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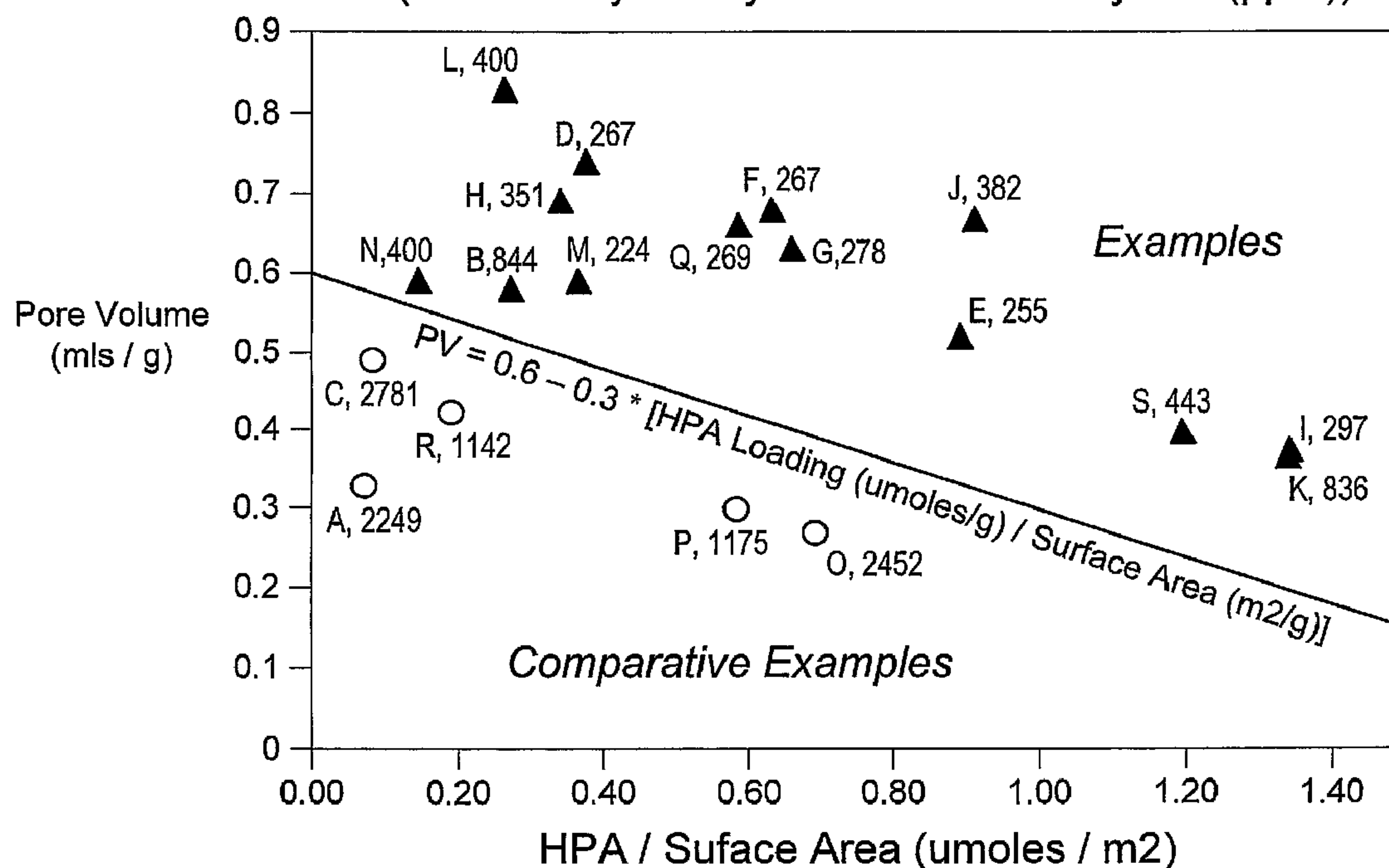
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(54) Titre : PROCEDE DE PRODUCTION D'ALCENES A PARTIR D'OXYGENATS EN UTILISANT DES CATALYSEURS
HETEROPOLYACIDES SUPPORTES
(54) Title: PROCESS FOR PRODUCING ALKENES FROM OXYGENATES BY USING SUPPORTED
HETEROPOLYACID CATALYSTS

Catalyst Pore Volume versus HPA Loading / Unit Surface Area
(Labelled by Catalyst & Ethane in Ethylene (ppm))



(57) Abrégé/Abstract:

The present invention relates to a supported heteropolyacid catalyst, to a process for producing alkenes from oxygenates in the presence of said catalyst, and to the use of said catalyst in a process for producing alkenes from oxygenates at a higher productivity whilst reducing the formation of alkanes.

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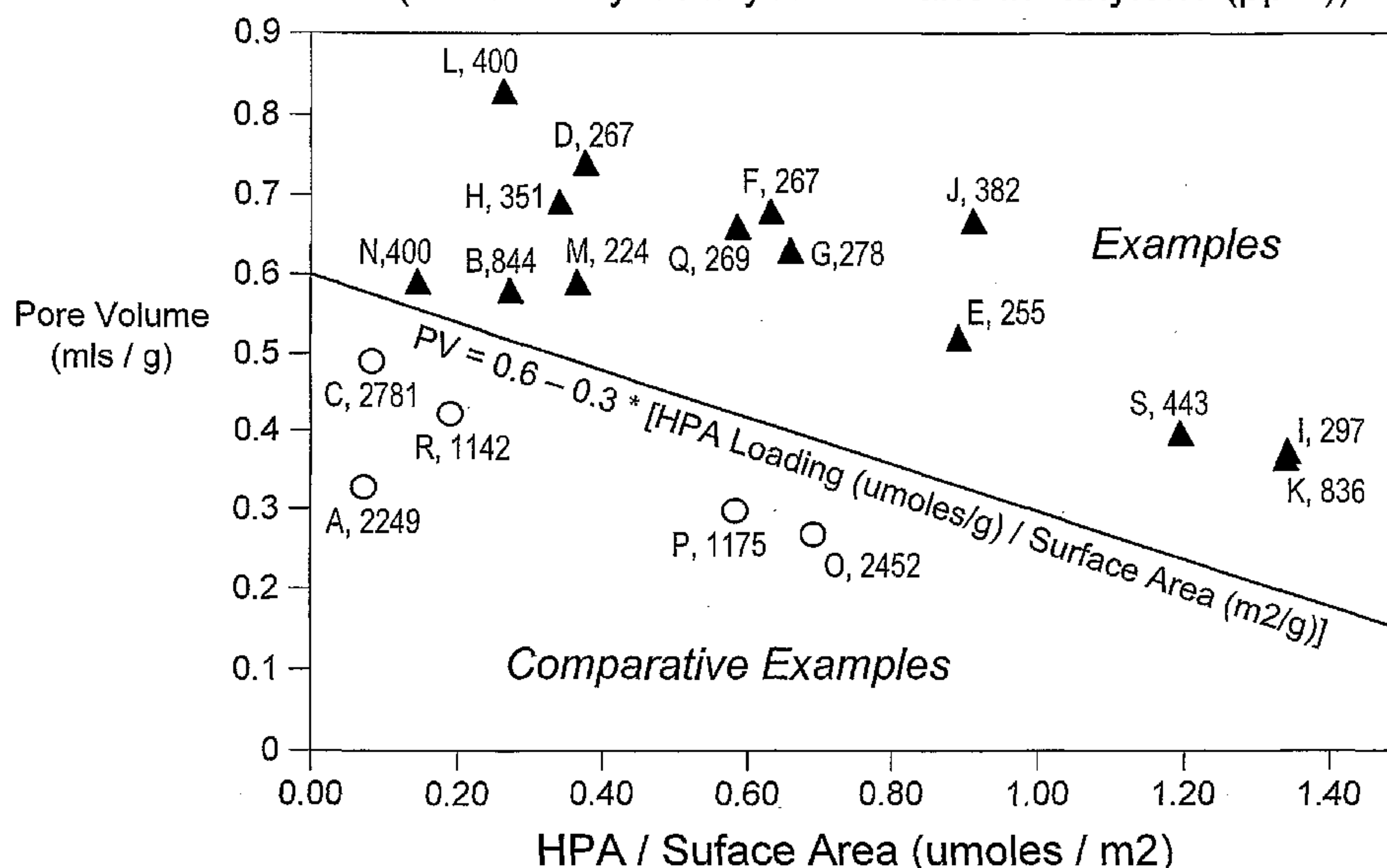
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(54) Title: PROCESS FOR PRODUCING ALKENES FROM OXYGENATES BY USING SUPPORTED HETEROPOLYACID CATALYSTS

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(57) Abstract: The present invention relates to a supported heteropolyacid catalyst, to a process for producing alkenes from oxygenates in the presence of said catalyst, and to the use of said catalyst in a process for producing alkenes from oxygenates at a higher productivity whilst reducing the formation of alkanes.

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PROCESS FOR PRODUCING ALKENES FROM OXYGENATES BY USING SUPPORTED
HETEROPOLYACID CATALYSTS

The present invention relates to a supported heteropolyacid catalyst and/or its salts
5 and to the production process of the said supported heteropolyacid catalyst and/or its salts.

In particular the present invention provides a use of a supported heteropolyacid
catalyst for converting oxygenates to alkenes, whereby an improved process is obtained, in
terms of alkene productivity and selectivity, whilst preventing the formation of alkanes.

The present invention further provides an improved process for the production of
10 alkene(s) from an oxygenate feedstock using a supported heteropolyacid catalyst under
specific operating conditions.

Heteropolyacids are valuable chemical compounds which can be used as acid
and/or oxygenate catalysts. The processes cited hereinafter, relate to a variety of examples
of various heteropolyacids and/or the method(s) for producing them;

15 US2006052240 describes a supported catalyst comprising a support having
supported thereon at least one member selected from the group consisting of
heteropolyacids and heteropolyacid salts, in which the heteropolyacid and/or
heteropolyacid salt is substantially present in a surface layer region of the support to a
depth of 30% from the support surface.

20 US6624325 describes a catalyst for producing lower fatty acid esters through
esterification of a lower aliphatic carboxylic acid with a lower olefin, comprising a
heteropoly acid or a salt thereof held on a carrier, and having a specific surface area of the
catalyst, as measured by a BET method, of 65 m²/g-350 m²/g. A process for
producing the catalyst and a process for producing a lower fatty acid ester by using the
25 catalyst are also provided.

US5227141 describes a membrane catalytic reactor which comprises a
heteropolyacid selected from the group consisting of 12-tungstophosphoric acid, 12-
molybdophosphoric acid, 12-molybdotungstophosphoric acid, and 12-tungstosilicic acid,
and polysulfone membrane is provided. This membrane catalytic reactor is applicable to
30 vapor-phase dehydration, dehydrogenation, oxidation, and simultaneous separation of
organic or inorganic materials, particularly vapor-phase dehydration of ethanol.

US2004/024918 describes a catalyst for use in producing a lower aliphatic

carboxylic acid ester, wherein the catalyst is produced by a process comprising a step of contacting the catalyst with a gas containing at least one member selected from water, lower aliphatic carboxylic acids and lower aliphatic alcohols; a process for producing the catalyst; and a process for producing a lower aliphatic carboxylic acid ester using the catalyst. The document further describes a siliceous support for use in a catalyst, which has a silicon content of from 39.7 to 46.3% by mass or a silicon content of from 85 to 99% by mass in terms of silicon dioxide or a crush strength of 30 N or more. By the use of a catalyst comprising the support, a lower aliphatic carboxylic acid ester is produced from lower olefin and a lower aliphatic carboxylic acid in a gas phase without causing great reduction of catalytic activity or cracking or abrasion of the catalyst.

Alkene(s) have traditionally been produced by steam or catalytic cracking of hydrocarbons. However, inevitably as oil resources are decreasing, the price of oil is increasing; making light alkene production a costly process. Thus there is an ever-growing need for non-petroleum routes to produce C₂ and C₂+ alkene(s), particularly ethene and propene, as these make useful starting materials for a number chemical processes, including the production of polymeric compounds, such as, polyethylene and polypropylene.

In recent years the search for alternative starting materials for C₂ and C₂+ alkene(s), production has led to the use of oxygenates such as alcohols (e.g. methanol, ethanol and higher alcohols). These alcohols provide a particularly attractive route for the production of alkene(s) since the said alcohols may be produced by the fermentation of, for example, sugars and/or cellulosic materials.

Alternatively, alcohols may be produced from synthesis gas. Synthesis gas refers to a combination of hydrogen and carbon oxides produced in a synthesis gas plant from a carbon source such as natural gas, petroleum liquids, biomass and carbonaceous materials including coal, recycled plastics, municipal wastes, or any organic material. Thus, alcohol and alcohol derivatives may provide non-petroleum based routes for the production of alkene(s) and other related hydrocarbons.

Generally, the production of alcohols, for example methanol, takes place via three process steps: synthesis gas preparation, methanol synthesis, and methanol purification. In the synthesis gas preparation step, an additional stage maybe employed whereby the feedstock is first treated, e.g. the feedstock is purified to remove sulphur and other

potential catalyst poisons prior to being converted into synthesis gas. This treatment can also be conducted after syngas preparation; for example, where coal or biomass is employed.

The processes cited hereinafter, relate to examples and alternative processes for the dehydration of an oxygenate feedstock; U.S. Pat. No. 4,543,435 discloses a process for converting an oxygenate feedstock comprising methanol, dimethyl ether or the like in an oxygenate conversion reactor into liquid hydrocarbons comprising C2 -C4 alkenes and C5+ hydrocarbons. The C2 -C4 alkenes are compressed to recover an ethylene-rich gas. The ethylene-rich gas is recycled to the oxygenate conversion reactor.

U.S. Pat. No. 4,076,761 discloses a process for converting oxygenates to gasoline with the return of a hydrogen-rich gaseous product to a synthesis gas plant or the oxygenate conversion reaction zone.

U.S. Pat. No. 5,177,114 discloses a process for the conversion of natural gas to gasoline grade liquid hydrocarbons and/or alkenes by converting the natural gas to a synthesis gas, and converting the synthesis gas to crude methanol and/or dimethyl ether and further converting the crude methanol/dimethyl ether to gasoline and alkenes. International Patent Application No. 93/13013 to Kvisle et al. relates to an improved method for producing a silicon-alumino-phosphate catalyst which is more stable to deactivation by coking. The patent discloses that after a period of time, all such catalysts used to convert methanol to olefin(s) (MTO) lose the active ability to convert methanol to hydrocarbons primarily because the microporous crystal structure is coked; that is, filled up with low volatility carbonaceous compounds which block the pore structure. The carbonaceous compounds can be removed by conventional methods such as combustion in air.

EPO publication No. 0 407 038A1 describes a method for producing dialkyl ethers comprising feeding a stream containing an alkyl alcohol to a distillation column reactor into a feed zone, contacting the stream with a fixed bed solid acidic catalytic distillation structure to form the corresponding dialkyl ether and water, and concurrently fractionating the ether product from the water and unreacted materials.

U.S. Pat. No 5,817,906 describes a process for producing light alkenes from a crude oxygenate feedstock comprising alcohol and water. The process employs two reaction stages. Firstly, the alcohol is converted using reaction with distillation to an ether. The

ether is then subsequently passed to an oxygenate conversion zone containing a metal aluminosilicate catalyst to produce a light olefin stream.

However, a major disadvantage associated with the typical dehydration methods used for producing alkene(s) from oxygenates, is that aromatic and alkane by-products are usually co-produced together with the targeted alkene products, and prove both difficult and expensive to separate, e.g. to recover ethene from ethane is both complex and expensive. For example, the well known methanol to olefin(s) - MTO- process can be described as the dehydrative coupling of methanol to alkenes and is a well known chemistry that can be employed to produce alkenes from alcohol(s). This mechanism is thought to proceed via a coupling of C1 fragments generated by the acid catalysed dehydration of methanol, possibly via a methyloxonium intermediate. Again, the main disadvantage of the said MTO process is that a range of alkenes are co-produced together with a range of various aromatic and alkane by-products, which in turn makes it very difficult and expensive to recover the desired alkenes.

Molecular sieves such as the microporous crystalline zeolite and non-zeolitic catalysts, particularly silicoaluminophosphates (SAPO), are known to promote the conversion of oxygenates by methanol to olefin (MTO) chemistry to hydrocarbon mixtures. Numerous patents describe this process for various types of these catalysts: U.S. Pat. Nos. 3,928,483, 4,025,575, 4,252,479 (Chang et al.); 4,496,786 (Santilli et al.); 4,547,616 (Avidan et al.); 4,677,243 (Kaiser); 4,843,183 (Inui); 4,499,314 (Seddon et al.); 4,447,669 (Harmon et al.); 5,095,163 (Barger); 5,191,141 (Barger); 5,126,308 (Barger); 4,973,792 (Lewis); and 4,861,938 (Lewis) .

Additionally, this reaction has a high activation energy step – possibly in the methanol or dimethylether production step - hence to achieve high conversion there is a need for high temperatures, e.g. 450°C, to drive the reactions forward. Conventionally, various means, such as, a heated catalyst recycle and/or “Downtherm” heating systems, have been implemented in such systems in order to obtain such high temperature conditions. Unfortunately operating at these said high temperatures leads to major problems, such as, catalyst deactivation, coking and by-product formation. In order to avoid these problems the reactions may be operated at lower temperatures, but this necessitates an expensive recycle of intermediates and reactants.

These and other disadvantages of the prior art show that there is a clear need,

within the field, for an improved and/or alternative process for the production of C2 and C2+ alkene(s) from oxygenates.

Hence, the present invention specifically sets to overcome these and other disadvantages and relates to

5 - a selected group of supported heteropolyacid catalyst and/or its salts, which has specific properties that allow it to have improved qualities when used in processes, such as, the dehydration of oxygenates to alkene(s),

 - the preparation of the said selected group of supported heteropolyacid catalyst and/or its salts,

10 - an improved process, in terms of selectivity, productivity and reduced alkane make, for the dehydration of an oxygenate feedstock in order to produce alkene(s), and to

 - the use of a a selected group of supported heteropolyacid catalyst and/or its salts in a process for the dehydration of oxygenates to alkene(s) for improving the alkene(s) selectivity and productivity and reducing the alkane make.

15 Thus, the present invention provides a supported heteropolyacid catalyst wherein the supported heteropolyacid catalyst is characterised in that the pore volume of the supported heteropolyacid catalyst suffices the following formula:

$$PV > 0.6 - 0.3 [HPA \text{ loading} / \text{Surface Area of Dried Catalyst}]$$

20 and the heteropolyacid loading, per surface area of the supported heteropolyacid catalyst is preferably more than 0.1 (micro moles/m²).

 The present invention also provides the use of a supported heteropolyacid catalyst in the process of producing alkene(s) from an oxygenate feedstock, for increasing the selectivity and productivity of alkenes and reducing the formation of alkanes, wherein the
25 supported heteropolyacid catalyst is characterised in that the pore volume of the supported heteropolyacid catalyst suffices the following formula:

$$PV > 0.6 - 0.3 [HPA \text{ loading} / \text{Surface Area of Dried Catalyst}]$$

30 The present invention further provides a process for the production of alkene(s) from an oxygenate feedstock in a reactor, characterised by performing the process at a

temperature preferably comprised between 180 and 250 °C, in the presence of a supported heteropolyacid catalyst; wherein the supported heteropolyacid catalyst is characterised in that the pore volume of the supported heteropolyacid catalyst suffices the following formula:

5

$$PV > 0.6 - 0.3 [\text{HPA loading} / \text{Surface Area of Dried Catalyst}]$$

Where for the purposes of the present invention and appending claims:

10

- **PV** is the Pore Volume of the Dried supported heteropolyacid Catalyst (measured in ml/g catalyst);
- **HPA loading** is the amount of heteropolyacid present in the Dried supported heteropolyacid Catalyst (measured in micro moles/g);
- **Surface Area of Dried Catalyst** is the surface area of the Dried supported heteropolyacid Catalyst (measured in m²/g).

15

The applicants believe that the production of the alkene(s) from the oxygenate feedstock, as stated by the present invention, is believed to proceed via a combination of two different dehydration mechanisms. The first being the direct dehydration of the oxygenates to the corresponding alkene(s) and water (as shown by Equation 1);

20

Equation 1



And the second being a two step mechanism, where the oxygenate feedstock is dehydrated via the production of an ether intermediate (as illustrated by Equations 2 and 3);

25

Equation 2



Equation 3



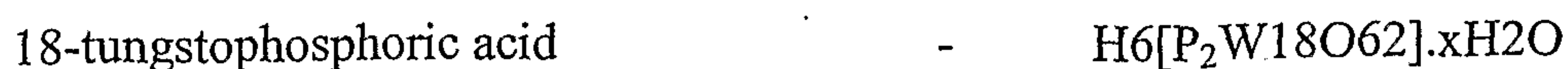
where R is an ethyl, propyl, butyl or pentyl group and R' is either hydrogen or a methyl, ethyl or propyl group. Preferably R is either a ethyl or propyl group and R' is either hydrogen or methyl moiety.

Equation 1, shows the endothermic direct elimination of alcohol to alkene(s) and water. Competing with Equation 1, are Equations 2 and 3; the exothermic etherification reaction (Equation 2), and the endothermic elimination of ether(s) to produce alkene(s) and alcohol (Equation 3). However, the overall dehydration of alcohols to alkene(s) is said to be an endothermic process.

According to the present invention and appending claims, all of the main reactions that are occurring (i.e. those reactions that are illustrated above) during the said dehydration process, are catalysed by the use of a supported heteropolyacid catalyst.

The term "heteropolyacid", as used herein and throughout the description of the present invention, is deemed to include inter alia; alkali, alkali earth, ammonium, free acids, bulky cation salts, and/or metal salts (where the salts may be either full or partial salts) of heteropolyacids. Hence, the heteropolyacids used in the present invention are complex, high molecular weight anions comprising oxygen-linked polyvalent metal atoms.

Typically, each anion comprises 12-18, oxygen-linked polyvalent metal atoms. The polyvalent metal atoms, known as the peripheral atoms, surround one or more of the central atoms in a symmetrical manner. The peripheral atoms may be one or more of molybdenum, tungsten, vanadium, niobium, tantalum, or any other polyvalent metal. The central atoms are preferably silicon or phosphorus, but may alternatively comprise any one of a large variety of atoms from Groups I-VIII in the Periodic Table of elements. These include copper, beryllium, zinc, cobalt, nickel, boron, aluminium, gallium, iron, cerium, arsenic, antimony, bismuth, chromium, rhodium, silicon, germanium, tin, titanium, zirconium, vanadium, sulphur, tellurium, manganese nickel, platinum, thorium, hafnium, cerium, arsenic, vanadium, antimony ions, tellurium and iodine. Suitable heteropolyacids include Keggin, Wells-Dawson and Anderson-Evans-Perloff heteropolyacids. Specific examples of suitable heteropolyacids are as follows:



	12-tungstophosphoric acid	-	H ₃ [PW ₁₂ O ₄₀].xH ₂ O
	12-molybdophosphoric acid	-	H ₃ [PMo ₁₂ O ₄₀].xH ₂ O
	12-tungstosilicic acid	-	H ₄ [SiW ₁₂ O ₄₀].xH ₂ O
	12-molybdosilicic acid	-	H ₄ [SiMo ₁₂ O ₄₀].xH ₂ O
5	Cesium hydrogen tungstosilicate	-	Cs ₃ H[SiW ₁₂ O ₄₀].xH ₂ O

and the free acid or partial salts of the following heteropolyacids:

	Monopotassium tungstophosphate	-	KH ₅ [P ₂ W ₁₈ O ₆₂].xH ₂ O
	Monosodium 12-tungstosilicic acid	-	NaK ₃ [SiW ₁₂ O ₄₀].xH ₂ O
	Potassium tungstophosphate	-	K ₆ [P ₂ W ₁₈ O ₆₂].xH ₂ O
10	Sodium molybdophosphate	-	Na ₃ [PMo ₁₂ O ₄₀].xH ₂ O
	Ammonium molybdodiphosphate	-	(NH ₄) ₆ [P ₂ Mo ₁₈ O ₆₂].xH ₂ O
	Potassium molybdodivanado phosphate	-	K ₅ [PMoV ₂ O ₄₀].xH ₂ O

Preferred examples of selected improved heteropolyacids are as follows:

	18-tungstophosphoric acid	-	H ₆ [P ₂ W ₁₈ O ₆₂].xH ₂ O
15	12-tungstophosphoric acid	-	H ₃ [PW ₁₂ O ₄₀].xH ₂ O
	12-tungstosilicic acid	-	H ₄ [SiW ₁₂ O ₄₀].xH ₂ O
	Cesium hydrogen tungstosilicate	-	Cs ₃ H[SiW ₁₂ O ₄₀].xH ₂ O

and the free acid or partial salts of the following heteropolyacids:

	Monopotassium tungstophosphate	-	KH ₅ [P ₂ W ₁₈ O ₆₂].xH ₂ O
20	Monosodium 12-tungstosilicic acid	-	NaK ₃ [SiW ₁₂ O ₄₀].xH ₂ O
	Potassium tungstophosphate	-	K ₆ [P ₂ W ₁₈ O ₆₂].xH ₂ O

In addition mixtures of different heteropolyacids and salts can be employed. The preferred heteropolyacids for use in the process described by the present invention is any one or more heteropolyacid that is based on the Keggin or Wells-Dawson structures; more preferably the chosen heteropolyacid for use in the process described by the present invention is any one or more of the following: silicotungstic acid, phosphotungstic acid, silicomolybdic acid and phosphomolybdic acid; and more preferably the chosen heteropolyacid for use in the process described by the present invention is any one or more silicotungstic acid; and most preferably the chosen heteropolyacid for use in the process described by the present invention is 12-tungstosilicic acid (H₄[SiW₁₂O₄₀].xH₂O)

Preferably, the heteropolyacids employed according to the present invention have molecular weights of more than 700 and less than 8500, preferably more than 2800 and

less than 6000. Such heteropolyacids also include dimeric complexes.

The supported heteropolyacid catalyst may be conveniently prepared by dissolving the chosen heteropolyacid in a suitable solvent, where suitable solvents include polar solvents such as water, ethers, alcohols, carboxylic acids, ketones and aldehydes and/or mixtures thereof and distilled water and/or ethanol are the most preferred solvents. The resulting acidic solution has a heteropolyacid concentration that is preferably comprised between 10 to 80 wt %, more preferably 20 to 70 wt% and most preferably 30 to 60 wt%. This said solution is then added to the chosen support (or alternatively the support is immersed in the solution). The actual volume of acidic solution added to the support is not restricted, and hence may be enough to achieve incipient wetness or wet impregnation, where wet impregnation (i.e. preparation using an excess acidic solution volume relative to pore volume of support), is the preferred method for the purposes of the present invention.

The resulting supported heteropolyacid may be modified, and various salts of heteropolyacid may then be formed in the aqueous solution either prior to, or during, impregnation of the acidic solution onto the support, by subjecting the supported heteropolyacid to a prolonged contact with a solution of a suitable metallic salt or by addition of phosphoric acid and/or other mineral acids.

When using a soluble metallic salt to modify the support, the salt is taken in the desired concentration, with the heteropolyacid solution. The support is then left to soak in the said acidic solution for a suitable duration (e.g. a few hours), with periodic stirring or shaking, after which time it is filtered, using suitable means, in order to remove any excess acid.

When the salt is insoluble it is preferred to impregnate the catalyst with the heteropolyacid and then to tritrate with the salt precursor. This method can improve the dispersion of the heteropolyacid salt. Other techniques such as vacuum impregnation may also be employed.

The impregnated support may then be washed and dried. This may be achieved using any conventional separation technique, including, for example, decantation and/or filtration. Once recovered, the impregnated support may be dried, preferably by placing the impregnated support in an oven at elevated temperature. Typically this drying is conducted at 130C with a nitrogen flow for 16 hours and then cooled to room temperature. For the purpose of the present description and appended claims, the term "Dried Catalyst" is

defined as the catalyst prepared by the preceding description. Alternatively, or additionally, a dessicator may be employed. On a commercial scale this drying stage is often achieved by a purge of hot inert gas such as nitrogen.

5 The amount of heteropolyacid impregnated onto the resulting support is suitably in the range of 10 wt % to 80 wt % and preferably in between 20 wt % to 50 wt %, based on the total weight of the heteropolyacid and of the support.

The weight of the catalyst on drying and the weight of the support used, may be used to obtain the weight of the acid on the support by deducting the latter from the former, giving the catalyst loading as a 'g heteropolyacid/kg catalyst' term. The catalyst loading in
10 'g heteropolyacid /litre support' can also be calculated by using the known or measured bulk density, of the support. The preferred catalytic loading of heteropolyacid is 150 to 600g heteropolyacid /kg catalyst

According to a preferred embodiment of the present invention the average heteropolyacid loading per surface area of the dried supported heteropolyacid catalyst is
15 more than 0.1 micro moles/m².

It should be noted that the polyvalent oxidation states and hydration states of the heteropolyacids stated previously, and as represented in the typical formulae of some specific compounds (shown above), only apply to the fresh acid before it is impregnated onto the support, and especially before it is subjected to the dehydration process conditions
20 of the present invention. The degree of hydration of the heteropolyacid may affect the acidity of the supported catalyst and hence its activity and selectivity. Thus, either or both of these actions of impregnation and dehydration process may change the hydration and oxidation state of the metals in the heteropolyacids, i.e. the actual catalytic species used, under the process conditions given, may not yield the hydration/oxidation states of the
25 metals in the heteropolyacids used to impregnate the support. Naturally therefore it is only to be expected that such hydration and oxidation states may also be different in the spent catalysts after the dehydration process of the present invention.

According to a preferred embodiment of the present invention, the amount of chloride present in/on the said heteropolyacid supported catalyst is less than 40 ppm,
30 preferably less than 25 ppm and most preferably less than 20 ppm.

Suitable catalyst supports may be in a powder form or alternatively may be in a granular form, or in a pelletised form, a spherical form or as a extrudates (including shaped

particles) and include, but are not limited to, Mordenites e.g. montmorillonite, clays, bentonite, diatomous earth, titania, activated carbon, alumina, silica-alumina, silica-titania cogels, silica-zirconia cogels, carbon coated alumina, zeolites, zinc oxide, flame pyrolysed oxides. Supports can be mixed oxides, neutral or weakly basic oxides. Silica supports are preferred, such as silica gel supports and supports produced by the flame hydrolysis of SiCl_4 . Preferred supports are substantially free of extraneous metals or elements which might adversely affect the catalytic activity of the system. Thus, suitable silica supports are at least 99% w/w pure. Impurities amount to less than 1% w/w, preferably less than 0.60% w/w and most preferably less than 0.30% w/w. The pore volume of the support is preferably more than 0.50ml/g and preferably more than 0.8 ml/g.

Suitable silica supports include, but are not limited to any of the following: Grace Davison Davicat® Grade 57, Grace Davison Davicat® 1252, Grace Davison Davicat® SI 1254, Fuji Silysia CariAct® Q15, Fuji Silysia CariAct® Q10, Degussa Aerolyst® 3045 and Degussa Aerolyst® 3043. The average diameter of the support particles is 2 to 10 mm, preferably 3 to 6 mm. However, these particles may be crushed and sieved to smaller sizes of, for example, 0.5-2 mm, if desired.

The average pore radius (prior to impregnation with the heteropolyacid) of the support is 10 to 500Å, preferably 30 to 175Å, more preferably 50 to 150 Å and most preferably 60 to 120Å. The BET surface area is preferably between 50 and 600 m²/g and is most preferably between 150 and 400 m²/g. The support has an average single particle crush strength of at least 1 kg force, suitably at least 2 kg force, preferably at least 6 kg force and more preferably at least 7 kg force. The bulk density of the support is at least 380 g/l, preferably at least 395 g/l.

The single particle crush strength was determined by using a Mecmesin force gauge which measures the minimum force necessary to crush a particle between parallel plates. The crush strength is based on the average of that determined for a set of at least 25 catalyst particles.

The BET surface area, pore volume, pore size distribution and average pore radius were determined from the nitrogen adsorption isotherm determined at 77K using a Micromeritics TRISTAR 3000 static volumetric adsorption analyser. The procedure used was an application of British Standard methods BS4359:Part 1:1984 'Recommendations for gas adsorption (BET) methods' and BS7591:Part 2:1992, 'Porosity and pore size

distribution of materials' - Method of evaluation by gas adsorption. The resulting data were reduced using the BET method (over the pressure range 0.05-0.20 P/Po) and the Barrett, Joyner & Halenda (BJH) method (for pore diameters of 20-1000 Å) to yield the surface area and pore size distribution respectively.

5 Suitable references for the above data reduction methods are Brunauer, S, Emmett, P H, & Teller, E, J. Amer. Chem. Soc. 60, 309, (1938) and Barrett, E P, Joyner, LG & Halenda P P, J. Am Chem. Soc. , 1951 73 373-380.

For the purpose of the above analytical measurements, samples of the supports and "Dried Catalysts" were out gassed for 16 hours at 120°C under a vacuum of 5×10^{-3} Torr.

10 A further embodiment of the said invention is where the chosen catalyst support is first treated with a fluorinating agent; the applicants believe that by fulfilling this said embodiment the catalyst will become more inert and/or acidic thus improving the selectivity and/or effectiveness of the catalyst during the aforementioned dehydration process.

15 According to the present invention, the feedstock that is used as the starting material for the dehydration reaction (as described hereinabove), is an "oxygenate feedstock" and preferably may comprise either of pure alcohol(s), or pure ether(s) or mixtures thereof.

20 Preferably the alcohol(s) present in the said oxygenate feedstock are least one monohydric aliphatic alcohols and have one or more alpha hydrogen(s) present, e.g. ethanol, n-propanol and t-butanol; which is a necessity in order to comply with the proposed mechanism of the invention by which the alkene(s) are produced (see Equations 1 to 3). The preferred monohydric aliphatic alcohol(s) are preferably ethanol and/or propanol, or a mixture thereof, for example, a mixture of ethanol and n-propanol and/or
25 iso-propanol. Ideally the alcohols present in the said oxygenate feedstock comprise a mixture of ethanol and n-propanol in order to be dehydrated into the corresponding "target alkene(s)". The term 'target alkenes', as used in the present description and appending claims hereinabove and hereinafter, is understood to mean the alkenes that are produced according to the process described by the present invention and preferably consist of either
30 ethene or propene or a mixture thereof.

According to a preferred embodiment of the present invention, the oxygenate feedstock to be dehydrated has an iso-propanol content of less than 5wt%, preferably less

than 1wt%, most preferably less than 0.1wt% and ideally contains no iso-propanol.

A preferred characterizing feature according to the present invention is that oxygenate feedstock to be dehydrated has a total C3+ alcohols (C3+ alcohols being defined as alcohols having at least 4 carbon atoms e.g. n-butanol, iso-butanol and pentanol) content of less than 5wt%, preferably less than 1wt%, most preferably less than 0.1wt% and ideally contains no C3+ alcohols.

Another preferred embodiment, according to the present invention, is that the alcohols present in the oxygenate feedstock to be dehydrated, have a methanol content of less than 5wt%, preferably less than 2 wt%, most preferably less than 0.5 wt% and ideally there is no methanol in the oxygenate feedstock. Corresponding advantages may accrue from eliminating methanol, i.e.

- (i) Prevention of dimethyl ether formation – dimethyl ether is hard to separate from propylene and ethylene compared to diethyl ether
- (ii) Prevention of MTO chemistry
- (iii) Prevention of alkylation of alkenes e.g. propylene to butene
- (iv) Prevention of formation of methyl ethyl ether (which is harder to separate from ethylene)
- (v) Less waste
- (vi) Lower toxicity
- (vii) Lower vapour pressure – easier to ship
- (viii) Shipping better C:O ratio in feedstock, i.e. less water production

Conventional distillation can be used according to the present invention in order to reduce/eliminate the methanol and C3+ alcohols from the alcohols present in the “oxygenate feedstock” to be dehydrated.

According to a preferred embodiment of the present invention, the ethers that may be used as a mixture with the alcohols present in the said “oxygenate feedstock” to be dehydrated, comprise any one or more of the homo, and mixed, ethers derived from the alcohols ethanol and propanol. For example diethyl ether, n-propyl ether, ethyl n-propyl ether, ethyl isopropyl ether, n-propyl isopropyl ether and iso-propyl ether.

The applicants have unexpectedly found that by using a mixture of the specified alcohol(s) and ether(s) (in terms of concentrations and components) as the oxygenate feedstock of the present invention, they were not only able to achieve a higher selectivity

towards the target alkenes, but they additionally found an overall increase in terms of productivity. Thus the present invention provides an improved process, in terms of reduced alkane make and increased selectivity and productivity of alkene(s), for the dehydration of a specific oxygenate feedstock to alkenes in the presence of a supported heteropolyacid catalyst.

According to a preferred embodiment of the present invention, water is permissible in the oxygenate feedstock to be dehydrated; and the said oxygenate feedstock may comprise up to 50 wt % of water, but preferably the said feedstock comprises less than 25 wt% water, and most preferably the feedstock comprises less than 20 wt % of water.

However due to processing costs such as the reactor size, heat of vaporization and heat capacity of water, it is preferred to operate with oxygenate feedstocks containing lower levels of water, for example less than 10wt%, preferably less than 5wt%.

As heteropolyacids are the chosen catalysts for the process, the levels of water in contact with the catalyst can affect the catalyst stability and activity. For example heteropolyacids show a diminished catalyst stability at low levels of water (<1wt%) and a diminished activity at high levels of water (>50wt%). To one skilled in the art it is apparent that the optimum water level will depend on the interaction of a complex set of variables including, alcohol feed composition, pressure, temperature and nature of the heteropolyacid employed. That said, this process has a good ability to separate water out and hence facilitates the use of bioethanol and other bioalcohol(s).

According to the present invention, the process whereby the oxygenate feedstock is dehydrated is preferably conducted in a dehydration reactor, with a vapour phase dehydration reactor being preferred. The reactor is engineered to cope with the exothermic ether formation and the endothermic dehydration to alkenes. The reaction temperature is preferably maintained within a small temperature range, as a temperature that is too low will reduce the rate of olefin manufacture and can lead to condensation of reactants; and too high a temperature can lead to the olefin being contaminated by unacceptable levels of by-products such as same carbon number alkanes. Preferably the temperature profile of the catalyst bed is less than 30°C, more preferably less than 15°C and most preferably less than 10°C. For a single bed adiabatic reactor, the overall endothermic reaction if allowed to go to thermodynamic equilibrium could result in a theoretical temperature drop of 180°C. Obviously the problem is one of heat management by reactor design. Suitable reactor

designs include those capable of handling heat fluxes such as fixed bed, fluidised bed, multi-tubular and multiple fixed bed reactors with inter-stage heaters. Optionally the heat management can be improved by injecting preheated fresh alcohol feed at several points in the reactor bed, at which point the exothermic etherification reaction can partially
5 counteract the overall endotherm. The feed can also be heated further, to above the reaction temperature, in order to provide an additional source of heat.

According to another preferred embodiment of the present invention the ethanol, propanol, ethers together with the water represent at least 90 wt % of the oxygenate feedstock introduced into the vapour phase dehydration reactor, preferably at least 99wt%
10 of the oxygenate feedstock introduced into the vapour phase dehydration reactor.

According to a preferred embodiment of the present invention, the operating conditions under which the dehydration process is conducted are such that the dehydration process is always operated in a vapour phase state. It is a preferred embodiment that the dehydration process operating pressure is always at least 0.1 MPa, preferably 0.2 MPa,
15 below the dew point pressure and/or that the dehydration process operating temperature is at least 10 °C above the dew point temperature of the feed entering the vapour phase dehydration reactor and the product composition that is present inside the vapour phase dehydration reactor. The latter being dependent on factors such as the initial feed composition and the degree of conversion within the reactor.

For the purposes of the present invention, the 'dew point temperature' is defined as being a threshold temperature. For example, for a given mixture, at a given pressure, if the system temperature is raised to above the dew point temperature, the mixture will exist as a dry gas. Likewise below the dew point temperature, the mixture will exist as a vapour containing some liquid. And similarly the 'dew point pressure', is defined as being a
25 threshold pressure. For example, for a given mixture, at a given temperature, if the system pressure is below the dew point pressure, the mixture will exist as a dry gas; above the dew point pressure, the mixture will exist as a vapour containing some liquid.

Thus, in keeping with the present invention, the specific temperature at which the dehydration process according to the present invention is conducted is preferably
30 comprised between 180 and 270 °C, more preferably between 190 and 260 °C and most preferably between 200 and 250 °C. The applicants have unexpectedly found that by operating at these said temperature ranges (which are particularly low when compared to

the typical temperatures employed in the field), they were not only able to benefit from the usual advantages of operating a process of a reduced temperature; but, they also found the process according to the present invention had the added advantage in that they were able to obtain a significantly reduced alkane make, e.g. < 1000 ppm of ethane in ethene.

5 The vapour phase reactor used for dehydrating the oxygenates is preferably operated at a pressure of above 0.1 MPa but less than 4.5 MPa, more preferably at a pressure of above 1.0 MPa but less than 3.5 MPa and most preferably at a pressure of above 1.0 MPa but less than 2.8 MPa.

The preferred reaction conditions used within the vapour phase dehydration reactor,
10 according to the present invention, are such that the dehydration process is always run at moderate conversion of oxygenates to corresponding alkenes.

For the purposes of the present invention and appending claims, moderate conversion of the "oxygenate feedstock" into corresponding alkenes, is defined as being the conversion of alcohols (e.g. ethanol and propanol) and/or their corresponding derived
15 ethers into the corresponding alkenes (e.g C2 and C3 alkenes) and means that 10 to 80%, more preferably 20 to 60%, of the alcohols and/or ethers are "converted" per pass. For a reactor with unreacted material recycle (ether and alcohol) operated in a manner to reach steady state the conversion approximates to:

The number of moles of alkenes produced per pass divided by the total number of
20 moles of alcohol and 2 times the ether in the feed stream containing the recycle and with added fresh feed.

This approximation hold true since the reaction is highly selective to alkenes under the preferred reaction conditions.

The experimental studies have shown that the present invention allows the
25 dehydration of an oxygenate feedstock into the corresponding alkenes to be conducted with a much higher selectivity than has previously been achieved in the field, e.g. current MTO chemistry. This said moderate conversion, low temperature, improved selectivity process benefits from improved economics as well as other advantages. Also, unlike conventional dehydration processes in the field, the present process described herein has no requirement
30 for a distinct and expensive alkane/alkene separation stage or catalyst regeneration.

The formation of ethers during the vapour phase dehydration process is thermodynamically favourable. This ether formation facilitates the separation of water

from the product stream as ethanol, n-propanol and iso-propanol are all fully or significantly water miscible and hence readily form water azeotropes during the dehydration process, which makes the separation of the water (a significant by-product of the reaction) from the product stream a very difficult process. However, the formation of
5 ethers such as diethyl ether and di-n-propyl ether, which both have a limited water miscibility and a very low water content azeotrope, allow the recovery of water by use of a decanter even in the presence of unreacted alcohols.

It should be noted that the operation of the process described by the present invention at medium conversion, with water removal during olefin recovery, has the
10 advantage in that it allows convergence towards the optimum reaction conditions with regards to the water content present in the reactor.

The liquid product stream after olefin removal comprises mostly unreacted alcohols, ethers and water. The applicants have found that it is preferable to recycle the major portion of the alcohols and ethers back to the vapour phase dehydration reactor after
15 water by-product removal. As indicated hereinabove, propanol can exist as two isomers n-propanol and iso-propanol; these isomers can interconvert under the reaction conditions so that the alcohol recycle stream may contain some iso-propanol in addition to unreacted ethanol and n-propanol. This isomerisation can also affect the compounds present in the ether proportion of the preferred recycle stream. The aforementioned ethers that are
20 optionally recycled back into the vapour phase dehydration reactor comprise some ethanol and propanol derived ether(s), such as, diethyl ether, n-propyl ether, ethyl n-propyl ether, ethyl isopropyl ether, n-propyl isopropyl ether and iso-propyl ether, which are produced during the dehydration stage.

According to a preferred embodiment of the present invention, a additional
25 separation stage is implemented into the process, whereby preferably at least 80 wt %; more preferably at least 90 wt %; most preferably at least 99 wt %; even more preferably at least 99.9 wt % of the ether(s) are separated from the alkenes. Thus, at least part, preferably all of the aforementioned ether(s) produced and recovered are then preferably recycled into the vapour phase dehydration reactor.

30 It is also an embodiment of the present invention that at least part, preferably all of the said ether recycle is pre-mixed with the oxygenate feedstock to be dehydrated, prior to entering the vapour phase dehydration reactor of the present invention.

The present invention also encompasses an embodiment whereby a process is provided for the conversion of hydrocarbon to alkene(s) comprising the consecutive steps of:

- a. converting a hydrocarbon feedstock, in a syngas reactor, into a mixture of carbon
oxide(s) and hydrogen,
- b. converting the said mixture of carbon oxide(s) and hydrogen from step a in the
presence of a particulate catalyst in a reactor under a temperature comprised
between 200 and 400°C and a pressure of 50 to 200 bar into a feedstock comprising
- at least one monohydric aliphatic paraffinic alcohol and/or corresponding ether -
having from 2 to 5 carbon atoms, and
- c. proceeding with the process, as described by the present invention, to produce
alkenes; whereby, the oxygenate feedstock comprises at least part of the alcohol(s)
and/or ethers produced in step b.

For the purpose of the above embodiment, any hydrocarbon containing feed stream that can be converted into a feedstock comprising carbon monoxide and hydrogen, most preferably a synthesis gas (or "syngas"), is useful in the processes of the invention.

The hydrocarbon feedstock used for syngas generation is preferably a carbonaceous material, for example biomass, plastic, naphtha, refinery bottoms, smelter off gas, municipal waste, coal, coke and/or natural gas, coal and natural gas being the preferred ones, most preferably natural gas.

Feedstocks comprising carbon monoxide and hydrogen, e.g., synthesis gas, may undergo purification prior to being fed to any reaction zones. Synthesis gas purification may be carried out by processes known in the art. See, for example, Weissermel, K. and Arpe H.-J., Industrial Organic Chemistry, Second, Revised and Extended Edition, 1993, pp. 19-21.

Examples and comparative examples

Support Materials used in the examples

- Silica pellets of CariAct® Q6, CariAct® Q10, CariAct® Q15, CariAct® Q30, CariAct® Q50 were obtained from Fuji Silysia

- Silica extrudates of Aerolyst® 3043 and Aerolyst® 3045 were obtained as extrudates from Degussa

- Silica granules of Davicat® Grade 57 and Davicat® 1252 and Davicat® 1254 were

obtained from Grace Davison

- Silica pellets Perlkat® 97-0, Perlkat® 39-3, Perlkat® 25-0 were obtained from Engelhard

- Silica aluminate pellets T4640 and T4641 were obtained from Sud Chemie

Support Properties

- 5 The support materials were analysed by nitrogen porosimetry (TABLE 1).

TABLE 1

Support	Surface Area (m ² /g)	Pore Volume (cm ³ /g)	Mean PSD (Å)
Perlkat 97-0	343	0.81	95
Perlkat 39-3	743	0.64	35
Davicat 1252	381	1.04	109
Davicat 1254	559	0.83	59
Aerolyst 3043	174	0.98	226
Aerolyst 3045	156	0.93	239
Sud Chemie T4640	156	0.37	94
Sud Chemie T4641	166	0.43	104
Perlkat 25-0	734	0.44	24
CariAct Q6	359	0.62	69
CariAct Q10	303	1.06	139
CariAct Q15	208	1.02	196
Cariact Q30	112	0.7	250
Cariact Q50	71	0.69	387
Davicat Grade 57	284	1.11	156

Catalyst Preparations using water as solvent

Silicotungstic acid ($\text{H}_4[\text{SiW}_{12}\text{O}_{40}]\cdot 24\text{H}_2\text{O}$, Mw 3310.6) was weighed into a wide-necked plastic bottle and dissolved in distilled water. To this acid solution was added a weighed amount of the support material. The acid solution and support material was left to
5 soak for approximately 1 hr with the occasional gentle swirl during this period to dislodge any trapped air bubbles.

After soaking the unadsorbed acid solution was removed by pouring the acid solution and support material out of the plastic container and into a plastic filter (which contained a filter paper).

10 The catalyst was allowed to drain until for approximately 15 to 60 minutes until no more liquid was being removed from the catalyst.

After draining was complete the catalyst was transferred to a ceramic tray and dried in a muffle furnace at 130°C under nitrogen

The dried solid catalyst was weighed and the amount of silicotungstic acid
15 adsorbed on the catalyst calculated by difference in weight to the starting material as indicated in the below TABLE 2.

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TABLE 2

Catalyst	Example	Support	Support wt (g)	HPA (g)	Water for HPA solution (g)	Weight catalyst (g)	HPA absorbed (g)	HPA Loading g/kg catalyst
A	Comparative	Perikat 25-0	30.09	15.49	47.52	34.25	4.16	121
B		Perikat 97-0	30.07	24.03	59.46	38.70	8.63	223
C	Comparative	Perikat 39-3	30.20	27.11	53.16	35.31	5.11	145
D		Davicat Grade G57	30.06	32.36	76.70	40.70	10.64	261
E		Davicat Grade G57	30.39	64.73	65.85	52.89	22.50	425
F		Aerolyst 3043	30.07	29.37	68.83	41.12	11.04	269
G		Aerolyst 3045	30.24	28.34	64.04	40.96	10.72	262
H		CariAct Q10	30.51	32.28	69.79	42.21	11.70	277
I		Davicat Grade G57	30.17	97.08	55.52	66.76	36.59	548
J		CariAct Q30	30.08	31.77	69.45	41.50	11.42	275
K		CariAct Q50	30.10	33.80	71.12	41.50	11.40	275
L		Davicat Grade G57	30.34	19.42	77.39	37.17	6.83	184
M		Davicat 1252	30.16	31.80	70.87	41.64	11.48	276
N		Davicat 1254	30.08	27.25	58.38	36.77	6.69	182
O	Comparative	T4640	30.17	31.77	54.07	37.92	7.76	205
P	Comparative	T4641	30.06	33.55	57.00	38.91	8.85	228
Q		CariAct Q15	15.04	15.61	34.56	21.16	6.12	289
R	Comparative	CariAct Q6	30.42	25.58	46.14	37.48	7.06	188

Catalyst Preparations using ethanol as solvent

The above procedure was followed except that ethanol was used as the solvent. A small amount of n-butanol was also added to the ethanol (TABLE 3).

TABLE 3

Catalyst	Example	Support	Support wt (g)	HPA (g)	EtOH & n-BuOH for HPA solution (g)	Weight catalyst g	HPA absorbed (g)	HPA Loading g/kg
S		Aerolyst 3043	30.59	62.87	47.969 & 1.499	59.63	29.04	487

Catalyst Properties

The dried solid catalysts were analysed by nitrogen porosimetry (TABLE 4).

TABLE 4

Catalyst	Example	Support	Surface Area of Catalyst (m ² /g)	Pore Volume of Catalyst (cm ³ /g Cat)	Mean PSD of Catalyst (Å)	HPA (g) / Surface Area of Catalyst (m ²)	HPA (μmoles) / Surface Area of Catalyst (m ²)	Calculation of Equation (1)
A	Comparative	Perkat 25-0	509	0.33	26	0.00024	0.07	0.5785
B		Perkat 97-0	246	0.58	95	0.00091	0.27	0.5178
C	Comparative	Perkat 39-3	518	0.49	38	0.00028	0.08	0.5746
D		Grace Davison G57	207	0.74	142	0.00126	0.38	0.4857
E		Grace Davison G57	144	0.52	144	0.00295	0.89	0.3325
F		Aerolyst 3043	128	0.68	212	0.00210	0.63	0.4095
G		Aerolyst 3045	120	0.63	211	0.00218	0.66	0.4021
H		CariAct Q10	242	0.69	114	0.00114	0.35	0.4963
I		Grace Davison G57	123	0.38	125	0.00446	1.35	0.1962
J		CariAct Q30	91	0.67	292	0.00302	0.91	0.3261
K		CariAct Q50	62	0.37	241	0.00444	1.34	0.1980
L		Grace Davison G57	209	0.83	159	0.00088	0.27	0.5202
M		Grace Davison 1252	227	0.59	104	0.00122	0.37	0.4898
N		Grace Davison 1254	374	0.59	64	0.00049	0.15	0.5559
O	Comparative	T4640	90	0.27	119	0.00228	0.69	0.3936
P	Comparative	T4641	118	0.3	103	0.00193	0.58	0.4249
Q		CariAct Q15	148	0.66	179	0.00195	0.59	0.4230
R	Comparative	CariAct Q6	298	0.42	57	0.00063	0.19	0.5428
S		Aerolyst 3043	123	0.4	129	0.00396	1.20	0.2411

Equation (1) = $0.6 - 0.3 \times [\text{HPA Loading } (\mu\text{moles/g}) / \text{Surface Area of Catalyst (m}^2/\text{g)}]$

When calculating the micromoles of heteropolyacid (HPA) adsorbed on the catalyst it is assumed that the heteropolyacid is fully hydrated. A Mw of 3310.6 was used for calculating the micro moles of 12-silicotungstic acid ($\text{H}_4[\text{SiW}_{12}\text{O}_{40}] \cdot 24\text{H}_2\text{O}$) adsorbed

Catalyst Testing

Catalysts (0.8 to 0.92ml, 125 to 180um particle size) were loaded into parallel flow reactors. Weights and volumes of the catalysts were recorded.

TABLE 5

Catalyst	Example	Support	Weight (g)	Volume (ml)
A	Comparative	Perlkat 25-0	0.6527	0.804
B		Perlkat 97-0	0.5365	0.859
C	Comparative	Perlkat 39-3	0.5001	0.901
D		Davicat G57	0.6531	0.845
E		Davicat G57	0.6438	0.707
F		Aerolyst 3043	0.5365	0.914
G		Aerolyst 3045	0.5365	0.914
H		CariAct Q10	0.5250	0.852
I		Davicat G57	0.7848	0.776
J		CariAct Q30	0.5120	0.887
K		CariAct Q50	0.5252	0.817
L		Davicat G57	0.4532	0.831
M		Davicat 1252	0.5439	0.776
N		Davicat 1254	0.5713	0.790
O	Comparative	T4640	0.6886	0.804
P	Comparative	T4641	0.6430	0.776
Q		CariAct Q15	0.5810	0.859
R	Comparative	CariAct Q6	0.7160	0.845
S		Aerolyst 3043	0.7623	0.804

The reactors were pressure tested and then heated to 220°C under a nitrogen flow. When the temperature had reached 220 °C the liquid feed of ethanol, diethyl ether and water were vapourised and mixed with nitrogen prior to being admitted to the reactor containing the

catalyst. The reaction conditions were 20 bar, nominal GHSV of 3000h⁻¹ for a 1ml catalyst bed, EtOH (28%v/v), Diethyl ether (34.5%v/v), water (3.3%v/v) and nitrogen (34.2%v/v).

When using small volume parallel flow reactors there will always be slight variations in the actual weight and volume of the catalyst(s) loaded into the separate reactor. Any

5 variation in the volume of the catalysts loaded then leads to a variation in the actual GHSV used since each individual reactor has the same total gas flow passing through it

The catalysts were tested under the sequence of temperatures; 220 °C for 24 hrs (Phase 1), 210 °C for 24hrs (Phase 2), 230 °C for 24hrs (Phase 3), 240 °C for 24 hrs (Phase 4) before finally being tested at 220°C for 24 hrs again (Phase 5).

10

TABLE 6

Catalyst	Example	Support	Reactor Temperature (°C)	reactor-pressure (bar)	GHSV (h ⁻¹)	Phase
A	Comparative	Perlkat 25-0	220	20.00	3697	Phase 5
B		Perlkat 97-0	220	19.97	3460	Phase 5
C	Comparative	Perlkat 39-3	220	19.97	3298	Phase 5
D		Davicat Grade 57	220	19.98	3517	Phase 5
E		Davicat Grade 57	220	20.01	4203	Phase 5
F		Aerolyst 3043	220	20.02	3252	Phase 1
G		Aerolyst 3045	220	19.97	3251	Phase 5
H		CariAct Q10	220	20.20	3488	Phase 5
I		Davicat Grade 57	220	20.01	3830	Phase 5
J		CariAct Q30	220	19.98	3350	Phase 5
K		CariAct Q50	220	20.02	3638	Phase 1
L		Davicat Grade 57	220	19.98	3577	Phase 5
M		Davicat 1252	220	20.01	3830	Phase 5
N		Davicat 1254	220	20.01	3762	Phase 5
O	Comparative	T4640	220	20.01	3696	Phase 5
P	Comparative	T4641	220	20.00	3829	Phase 5
Q		CariAct Q15	220	20.20	3460	Phase 5
R	Comparative	CariAct Q6	220	20.20	3518	Phase 5
S		Aerolyst 3043	220	20.01	3697	Phase 5

The flow rate of the exit gas from the reactors was measured and the exit gas was also

15 analysed by a GC using flame ionisation detectors. Ethylene, ethane, ethanol, diethyl ether,

butanes, butanes and acetaldehyde were detected by GC. Product rates were calculated from the gas flow and analysed composition.

The catalysts exhibited the following average catalyst performance, over the phase period defined above, as regards ethylene productivity and ethane selectivity (as defined by ethane / ethylene ratio - (TABLE 7).

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TABLE 7

Catalyst	Example	Support	Pore Volume of Catalyst (cm ³ /g Cat)	HPA (umoles) / Surface Area of Catalyst (m ²)	Calculation of Equation (1)	HPA g/kg cat	Ethene (g/kg cat/h)	Ethane (g/kg cat/h)	Ethane in Ethylene (ppm)
A	Comparative	Perikat 25-0	0.33	0.07	0.58	121	295	0.66	2249
B		Perikat 97-0	0.58	0.27	0.52	223	1222	1.03	844
C	Comparative	Perikat 39-3	0.49	0.08	0.57	145	59	0.17	2781
D		Grace Davison G57	0.74	0.38	0.49	261	1676	0.45	267
E		Grace Davison G57	0.52	0.89	0.33	425	2095	0.53	255
F		Aerolyst 3043	0.68	0.63	0.41	262	1782	0.72	406
G		Aerolyst 3045	0.63	0.66	0.40	262	1966	0.55	278
H		CariAct Q10	0.69	0.35	0.50	277	1622	0.57	351
I		Grace Davison G57	0.38	1.35	0.20	548	1636	0.49	297
J		CariAct Q30	0.67	0.91	0.33	275	1502	0.57	382
K		CariAct Q50	0.37	1.34	0.20	275	1234	1.04	836
L		Grace Davison G57	0.83	0.27	0.52	184	1343	0.54	400
M		Grace Davison 1252	0.59	0.37	0.49	276	1945	0.43	224
N		Grace Davison 1254	0.59	0.15	0.56	182	1398	0.56	400
O	Comparative	T4640	0.27	0.69	0.39	205	461	1.13	2452
P	Comparative	T4641	0.3	0.58	0.42	228	743	0.87	1175
Q		CariAct Q15	0.66	0.59	0.42	289	1706	0.46	269
R	Comparative	CariAct Q6	0.422	0.19	0.54	188	878	1.00	1142
S		Aerolyst 3043	0.4	1.20	0.24	487	1897	0.84	443

Equation (1) = $0.6 - 0.3 \times [\text{HPA Loading (umoles/g)} / \text{Surface Area of Catalyst (m}^2\text{/g)}]$

The mass balance was consistently above 95%, the detected by-products are given as ratios to the ethene, the analytical system employed was capable of detecting C1 – C6 hydrocarbons and oxygenates. As illustrated above the catalysts according to the present invention are highly selective.

- 5 These examples and comparative examples have been plotted into figure 1 which represents the catalyst pore volume (cm³/g) versus the ratio of HPA loading (umoles/g) to surface area (m²/g) of the catalyst. The value of the ethane in ethene (ppm) has been indicated on each example in the figure.

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Claims as amended under Art. 34 of the PCT, 17 September 2008:

1. A process for the production of alkene(s) from an oxygenate feedstock, in a reactor, in the presence of a supported heteropolyacid catalyst, characterised in that the pore
5 volume of the supported heteropolyacid catalyst suffices the following formula:

$$PV > 0.6 - 0.3 [\text{HPA loading} / \text{Surface Area of Dried Catalyst}]$$

where;

- PV is the Pore Volume of the dried supported heteropolyacid catalyst (measured in ml/g catalyst);
 - 10 • **HPA loading** is the amount of heteropolyacid present in the dried supported heteropolyacid catalyst (measured in micro moles/g);
 - **Surface Area of Dried Catalyst** is the surface area of the dried supported heteropolyacid catalyst (measured in m²/g).
2. A process according to Claim 1 wherein the heteropolyacid loading, per surface
15 area of the supported heteropolyacid catalyst is more than 0.1 micro moles/m².
3. A process according to any of the preceding claims, wherein the supported heteropolyacid catalyst has a catalytic loading of between 150 to 600g heteropolyacid /kg catalyst.
4. A process according to any of the preceding claims, wherein the pore volume of
20 the catalyst support of the supported heteropolyacid catalyst is preferably more than 0.50ml/g and preferably more than 0.8 ml/g.
5. A process according to any of the preceding claims, wherein the catalyst support of the supported heteropolyacid catalyst has an average pore radius of between 10 to 500Å, preferably between 30 to 175Å, more preferably between 50 to 150 Å and most preferably
25 between 60 to 120Å.
6. A process according to any of the preceding claims, wherein the catalyst support of the supported heteropolyacid catalyst has a BET surface area between 50 and 600 m²/g and preferably between 150 and 400 m²/g.
7. A process according to any of the preceding claims, wherein the catalyst support of
30 the supported heteropolyacid catalyst has an average single particle crush strength of at least 1 kg force, preferably at least 2 kg force, more preferably at least 6 kg force and most preferably at least 7 kg force.

8. A process according to any of the preceding claims, wherein the catalyst support of the supported heteropolyacid catalyst has a bulk density of at least 380 g/l and preferably at least 395 g/l.
9. A process according to any of the preceding claims, wherein the catalyst support of the supported heteropolyacid catalyst is a silica support.
10. A process according to any of the preceding claims, wherein the process is performed at a temperature comprised between 180 and 250 °C.
11. A process according to any of the preceding claims, whereby the dehydration process is conducted in the vapour phase state.
12. A process according to any of the preceding claims, whereby the dehydration process is operated at a pressure of at least 0.1 MPa, preferably at least 0.2 MPa, below the dew point pressure and/or that the dehydration process is operated at a temperature of at least 10 °C above the dew point temperature of the feed entering the vapour phase dehydration reactor and the product composition that is present inside the vapour phase dehydration reactor.
13. A process according to any of the preceding claims, whereby the dehydration process is operated at a temperature of between 180 and 270 °C, more preferably between 190 and 260 °C and most preferably between 200 and 250 °C.
14. A process according to any of the preceding claims, whereby the dehydration process is operated at a pressure of above 0.1 MPa but less than 4.5 MPa, more preferably at a pressure of above 1.0 MPa but less than 3.5 MPa and most preferably at a pressure of above 1.0 MPa but less than 2.8 MPa.
15. A process according to any of the preceding claims whereby the majority of the unreacted alcohols and/or ethers produced during the dehydration process are recycled back to the vapour phase dehydration reactor.
16. A process according to Claim 15, where at least part, preferably all of the said unreacted alcohols and/or ether recycle is pre-mixed with the oxygenate feedstock to be dehydrated prior to entering the vapour phase dehydration reactor.
17. A process for the conversion of hydrocarbon to alkene(s) comprising the consecutive steps of:
- a) converting a hydrocarbon feedstock, in a syngas reactor, into a mixture of carbon oxide(s) and hydrogen,

b) converting the said mixture of carbon oxide(s) and hydrogen from step a in the presence of a particulate catalyst in a reactor under a temperature comprised between 200 and 400°C and a pressure of 50 to 200 bar into a feedstock comprising - at least one monohydric aliphatic paraffinic alcohol and/or corresponding ether - having from 2 to 5 carbon atoms, and

c) proceeding with the process, as described by any of the preceding claims, to produce alkenes; whereby, the oxygenate feedstock comprises at least part of the alcohol(s) and/or ethers produced in step b.

18. Use of a supported heteropolyacid catalyst in a process of producing alkene(s) from an oxygenate feedstock for increasing the selectivity and productivity of alkenes, whilst preventing the formation of alkanes; wherein the pore volume of the supported heteropolyacid catalyst suffices the following formula:

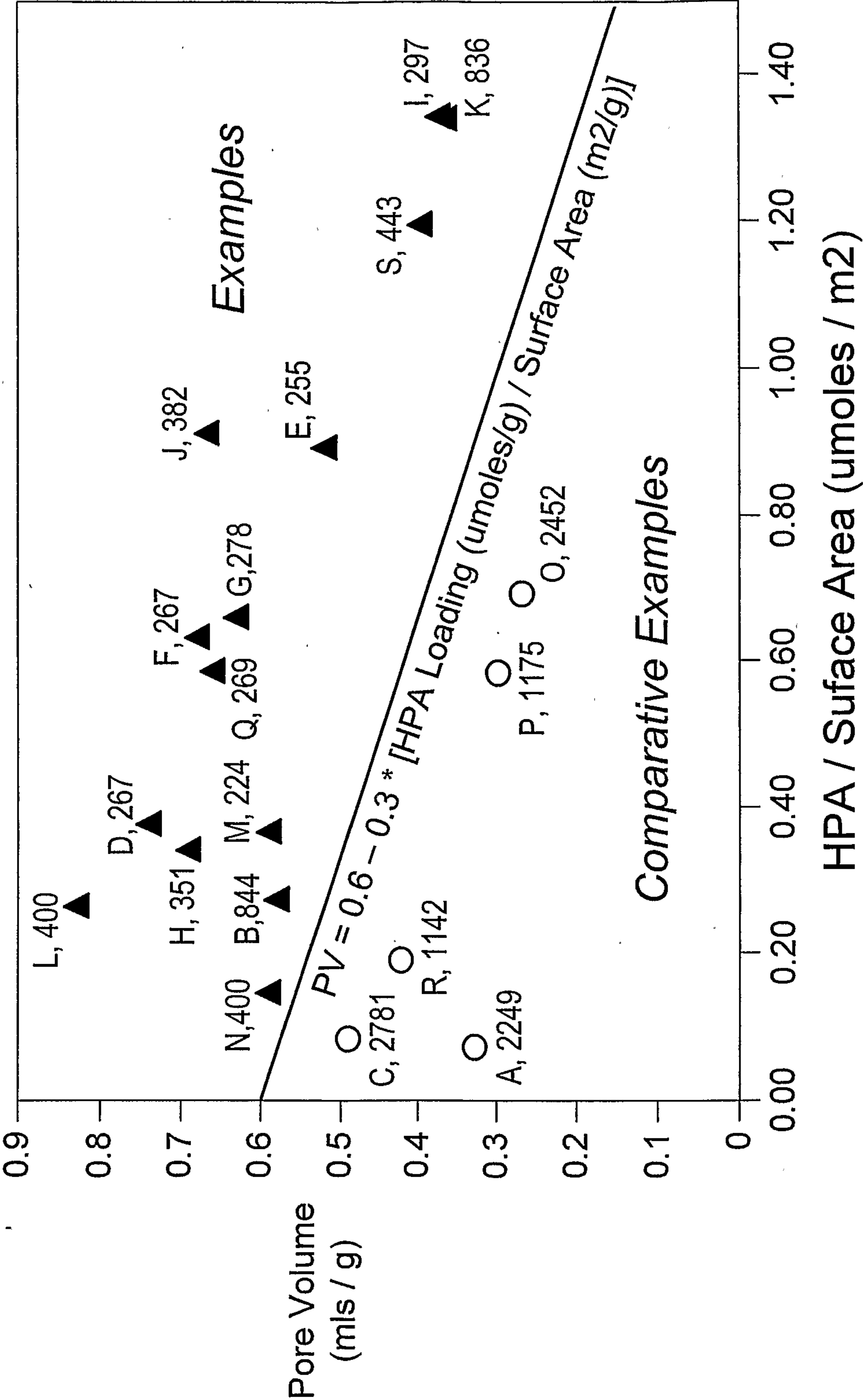
PV > 0.6 - 0.3 [HPA loading / Surface Area of Dried Catalyst].

19. Use according to claim 18, wherein the process of producing alkene(s) is according to any of claims 2 to 16.

20. Use of a supported heteropolyacid catalyst in step c) of the process according to claim 17, for increasing the selectivity and productivity of alkenes, whilst preventing the formation of alkanes; wherein the pore volume of the supported heteropolyacid catalyst suffices the following formula:

PV > 0.6 - 0.3 [HPA loading / Surface Area of Dried Catalyst].

Catalyst Pore Volume versus HPA Loading / Unit Surface Area
(Labelled by Catalyst & Ethane in Ethylene (ppm))



Comparative Examples

Catalyst Pore Volume versus HPA Loading / Unit Surface Area (Labelled by Catalyst & Ethane in Ethylene (ppm))

