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(54) BARIUM OXIDE COATED ZIRCONIA PARTICLE FOR USE IN AN
 OXYGEN GENERATOR

(71) We, THE BENDIX CORPORATION, a corporation organized and existing under the laws of the State of Delaware, United States of America, of Executive Offices, 5 Bendix Center, Southfield, Michigan, United States of America, do hereby declare the invention, for which we pray that a patent may be granted to us, and the method by which it is to be performed, to be particularly described in and by the following statement:—

The present invention relates to a process of manufacturing an oxide material for use in an oxygen generator, and the 10 invention also covers any oxide material produced with such a process.

For a time, oxygen was commercially produced by the Brin process which is 15 completely disclosed in U.S. Patent No. 432,815. In the Brin process, beds of barium oxide particles are heated in a chamber to a temperature of between 700°C—750°C. This temperature is maintained substantially constant and clean, dry air under pressure is passed through the chamber. The 20 barium oxide reacts with the oxygen in the air to form barium peroxide while nitrogen is released as a by-product. After a period of time, the flow of the pressurized air is interrupted and the chamber is connected to a source of vacuum. As 25 the gas in the chamber is evacuated, the barium peroxide releases the earlier acquired oxygen as it reverts back to barium oxide. The released oxygen is captured and communicated to a storage container for later 30 use.

From experimentation, it was determined 35 that the efficiency of the Brin process could be increased if the temperature of the barium oxide particles could be maintained at a constant temperature. The reaction between oxygen in the air and the barium oxide increases with temperature 40 and reaches a maximum near the melting point of the barium oxide. Unfortunately, maintaining the temperature of the chamber at a constant temperature requires sophisticated heaters to prevent fluctuation 45 in the chamber temperature since the barium

oxide particles are alternately being subjected to air under pressure from outside the chamber and vacuum and also since the heat of reaction removes heat from and supplies heat to the barium oxide particles. In addition, the barium oxide particles located nearer the heaters receive the greatest heat while the barium oxide particles away from the heater are relatively cooler. In an effort to heat the barium oxide particles farthest away from the heaters, the temperature of the barium oxide particles adjacent the heater approach the melting temperature. Unfortunately, the heat generated through the reaction of air and the barium oxide and the chamber temperature is additive, and at times sufficient to cause the barium oxide particles to melt and fuse into a larger mass. This is not too serious a problem during the reaction part of the process since the pressure of the air flowing through the chamber can be increased to maintain uniform flow through the bed. However, channels and cracks form in the bed of barium oxide particles during the oxygen evacuation part of the cycle resulting in deterioration of the barium oxide beds since the air flow thereafter is not uniform.

In an effort to sustain the efficiency of barium oxide-barium peroxide in an oxygen generator, the present invention proposes a composition of material which absorbs the heat of reaction created between oxygen in the air and barium oxide to thereby maintain the operating temperature below the melting temperature of the barium oxide.

It is thus an object of this invention to provide an oxygen generator with a composition of material which has heat capacity to absorb the heat caused by reacting barium oxide with oxygen in pressurized air to stabilize the reaction temperature below the melting point of barium oxide to provide a faster and more complete reaction. This composition of material should, in addition, have structural strength to sustain vigorous pressure cycling at elevated temperatures.

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It is another object of the invention to provide a process for manufacturing such a composition of material for use in a barium oxide-barium peroxide oxygen generator.

5 According to the invention there is provided an oxide composition for use in an oxygen generator having an optimum operating range of between 750°C—850°C, wherein said oxide composition reacts with 10 pressurized air to form a higher oxide and said higher oxide reverts to oxide through the release of oxygen in the absence of pressurized air, said oxide composition comprising barium oxide coated zirconia 15 particles.

The invention also provides a process of manufacturing an oxide material for use in an oxygen generator, comprising the steps of mixing a quantity of dry zirconia 20 particles with a quantity of barium peroxide particles to create a blended mixture, heating said blended mixture until said barium peroxide particles melt and coat the zirconia particles, decomposing said 25 peroxide coating to barium oxide, cooling the barium oxide coated zirconia particles and passing the cooled barium oxide coated zirconia particles through a grinder to produce a desired particle size for said oxide 30 material.

When these barium oxide coated zirconia particles are used in the oxygen generator, the barium oxide reacts with the oxygen in the pressurized air to create barium peroxide while producing nitrogen as a by-product. The heat created by the reaction of oxygen and barium oxide is carried into the zirconia particle through conduction. Any increase in the temperature in the 35 chamber during the oxidation portion of the cycle is dissipated through the evacuation of air from the chamber during the oxygen release portion of the cycle. In addition the heat absorbed into the zirconia particles is returned to the coating during the 40 release of oxygen. Thus, the temperature in the chamber is maintained within a set range below the melting point of barium oxide during an operating cycle of the oxygen generator.

45 The inventive concept will be explained in further detail in the following description of a preferred embodiment, given by way of example only and with reference to the 50 accompanying drawings, in which:

55 Figure 1 is a flow chart showing a method of manufacturing an oxide composition of material according to the invention; and

60 Figure 2 is a graph showing the oxygen yield vs. operation time for a number of operation cycles.

65 Referring now to the drawings, Figure 1 illustrates a process for making a com-

position of material for use in an oxygen generator according to the invention.

65 The composition of material consists of a zirconia particle having a coating of barium oxide located on its periphery.

70 This composition of material is manufactured by the following steps:

A quantity of zirconia particles having a size preferably between 0.59 and 0.25mm (30—60 mesh in US standards) is placed in a container and heat treated in an oven, 75 illustrated by step 1 in Figure 1. The temperature in the oven is raised to between 800—900°C and maintained at this temperature for about one hour. This heat treatment allows the zirconia to expand and relieve any internal stresses which may have occurred during the manufacture of the zirconia.

80 Thereafter, as shown by step 2, the temperature in the oven is uniformly reduced to allow the zirconia particles to cool to between 100—150°C. The zirconia particles are then removed from the oven and placed in a mixer.

85 As illustrated by step 3, in Figure 1, barium peroxide particles are added to the zirconia particles in the mixer and blended therewith until a uniform mixture is obtained.

90 This uniform mixture is placed in a container and returned to the oven, as 95 illustrated by step 4 in Figure 1. The air in the oven is replaced with nitrogen and the temperature therein is raised to between 500—600°C. This temperature is 100 maintained for about 1 hour or until any water or moisture which may be in the mixture has completely evaporated. At this temperature, the barium peroxide melts and flows around the zirconia particles. 105 Depending upon the size of the container, it may be necessary under some circumstances to stir the mixture to assure that all the zirconia particles are coated with the liquid barium peroxide. Thereafter, 110 the temperature of the oven is uniformly increased to 850°C causing the barium peroxide coating to decompose to barium oxide. The barium oxide remains relatively stable 115 since it does not react with nitrogen. However, it is speculated that a layer of barium zirconate is created along the interface between the periphery of the zirconia particle and the barium oxide. To assure that this stable condition is maintained, nitrogen 120 gas continually flows through the oven. After a period of time, typically about 4 hours, the temperature in the oven is uniformly reduced and, as illustrated by step 5 in Figure 1, the barium oxide coated zirconia particles are cooled to a temperature which may be about 40°C or room temperature. During this cooling step, the flow of nitrogen gas continues to maintain the

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stability of the barium oxide. When the barium oxide cools, it forms a matrix which holds large quantities of the zirconia particles together. Therefore, in order to obtain a uniform particle size the entire quantity of barium oxide coated zirconia particles is passed through a grinder, as illustrated by step 6 in Figure 1. As the ground barium oxide coated zirconia particles pass from the grinder they are carried to a mechanical sieve shaker where any residual barium oxide which has not bonded to the zirconia particle is removed. Thereafter, a uniformly sized barium oxide coated zirconia particle is removed and stored for later use in an oxygen generator. The ratio of the barium oxide coating to the zirconia particles varies in a range of 1% to 44% by weight.

In order to evaluate oxide composition made according to the manufacturing process described in Figure 1, the relationship of the barium oxide and zirconia particles were varied in the following examples:

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EXAMPLE #1

Zirconia particles having a 0.42mm (40 mesh in US standards) size were heat treated by heating to 850°C for 1 hour in an oven. These heat treated zirconia particles were cooled and placed in a blender and barium peroxide particles were added until a uniform mix having a ratio of about 2:1 was achieved. This uniform mixture was heated in an oven having a temperature of 600°C through which nitrogen gas was flowing at a rate of about 1 liter/min. After about an hour, any moisture contained in the mixture was evaporated, and barium peroxide melted and dispersed around zirconia particles. The temperature of the oven was increased to 850°C. When the temperature of the mixture reaches 850°C, the barium peroxide coating on the zirconia particles decomposes to barium oxide. At the end of 4 hours, the zirconia particles were coated with a barium oxide layer. Thereafter, the temperature of the oven was uniformly reduced and the mixture allowed to cool to about 40°C or room temperature. During this entire cooling procedure, the flow of nitrogen gas through the oven is continued to prevent the barium oxide from reacting with air which could be communicated to the oven. The mixture has an appearance of a large mass since the barium oxide forms a matrix during the cooling. This large mass of material is placed in a grinder having a set particle size of about 0.35mm (45 mesh in US Standards). The ground material is placed on a sieve and shaken in a mechanical shaker to remove any residual barium oxide which is not bonded to the zirconia particles. After about 15 minutes the desired

size barium oxide coated zirconia particles are essentially uniform in appearance and shape. From a quantitative analysis, it was determined that the amount of barium oxide coating on the zirconia particle was about 6% of the total weight of the coated zirconia particle.

EXAMPLE #2

The same particle size and amount of zirconia used as in Example #1, was used in Example #2, however, the amount of barium oxide was increased in the blended mixture to about 3:1. Thereafter, the manufacturing procedure as described in Example #1 was followed. However, the particle size of the grinder was decreased to about the 0.297mm (50 mesh size in US Standards). After passing through the sieve and shaker apparatus, a barium oxide coated zirconia particle having a uniform size and appearance was produced. From a quantitative analysis, it was determined that the barium oxide coating on the zirconia particle accounted for about 20% by weight of the total weight of the coated zirconia particle.

EXAMPLE #3

In order to investigate the optimum barium oxide coating the same particle size and amount of zirconia as illustrated in Example #1 was used and the amount of barium oxide increased in the blended mixture to about 4:1. Thereafter, the same manufacturing procedure as described in Example #1 was used followed until the grinding. The particle size of the grinder was increased to about the 60 mesh size. After passing through the sieve and shaker apparatus, a barium oxide coated zirconia particle having a uniform size and appearance was produced. From a quantitative analysis, it was determined that the barium oxide coating on the zirconia particle accounted for about 40% by weight of the total weight of the coated zirconia particle.

In order to evaluate the oxygen extraction capability of the composition of materials produced in Examples 1, 2 and 3, a quantity of each example was placed in an oxygen extractor apparatus similar to that disclosed in U. S. Patent No. 3,720,501.

The graph of Figure 2 illustrates the effectiveness of the materials produced in Examples 1, 2 and 3 as compared to the effectiveness of barium oxide.

As seen in the graph, after 50 hours of continuous operation, the oxygen extraction capability of a bed of pure barium oxide particles deteriorates due to excessive amounts of heat generated during the reaction of oxygen in the input air with the barium oxide to produce barium peroxide.

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Once a bed of barium oxide particles start to cake or fuse together, the oxygen output falls off to a point where at about 70 hours the production is about .05 kg of oxygen per hour/kg bed material.

From the graph it should be evident that the most effective material for continuous oxygen production is that produced in Example 1. It is speculated that the 6% barium oxide by weight of the total mixture is just sufficient to coat the peripheral surface area of the zirconia particle with a thin layer of barium oxide. The heat generated by the surface reaction of the barium oxide with oxygen in the air is carried into the zirconia particle to maintain the temperature within the chamber below the melting point of the barium oxide. As illustrated in Figure 2, this type of composition is capable of operating over 300 hours without a decrease in oxygen production.

The composition of material in Example 2, where the barium oxide is about 20% of the total weight of the bed material while providing continuous performance similar to Example 1 only produced about one-half of the oxygen output per unit of time.

The composition of material in Example 3, where the barium oxide is about 40% of the total weight of the bed material, while providing for continuous performance about twice as long as the barium oxide reference only produced about one-third the amount of oxygen per unit of time.

From these tests it was determined that: a composition of material made up of barium oxide coated zirconia particle could be used to extract oxygen from the air without deteriorating when continually used for a time period over 300 hours; the operational life of the composition of material decreased as the ratio of barium oxide to zirconia increased; and the zirconia particle has sufficient heat absorbing capability to stabilize the reaction temperature between barium oxide and oxygen in the air below the melting point of barium oxide to maintain the temperature in a reaction chamber at a substantially uniform temperature.

WHAT WE CLAIM IS:—

- An oxide composition for use in an oxygen generator having an optimum operating range of between 750—850°C, wherein said oxide composition reacts with pressurized air to form a higher oxide and said higher oxide reverts to oxide through the release of oxygen in the absence of pressurized air, said oxide composition comprising barium oxide coated zirconia particles.
- An oxide composition according to Claim 1, wherein the ratio of the barium oxide coating to the zirconia particles varies in a range of 1% to 44% by weight.
- An oxide composition according to either preceding claim, wherein the zirconia particle size varies from 0.59 to 0.25 mm (30 to 60 mesh in US standards).
- A process of manufacturing an oxide material for use in an oxygen generator, comprising the steps of mixing a quantity of dry zirconia particles with a quantity of barium peroxide particles to create a blended mixture, heating said blended mixture until said barium peroxide particles melt and coat the zirconia particles, decomposing said peroxide coating to barium oxide, cooling the barium oxide coated zirconia particles and passing the cooled barium oxide coated zirconia particles through a grinder to produce a desired particle size for said oxide material.
- A process according to Claim 4, further including the step of purging said blended mixture with nitrogen gas during said heating to maintain the barium oxide in a stable condition.
- A process according to Claim 5, further including the step of purging said barium oxide coated zirconia particles with nitrogen gas during said cooling to maintain the stable formation of oxide.
- A process according to any of Claims 4 to 6, further including the step of maintaining the blended mixture at a temperature below the melting temperature of said barium peroxide for a time period sufficient to evaporate any water contained therein which could affect the coating of the zirconia particles by the barium oxide particles.
- A process of manufacturing an oxide material for use in an oxygen generator, substantially as described hereinabove with reference to the accompanying drawings.
- An oxide composition for use in an oxygen generator, substantially as described hereinabove with reference to the accompanying drawings.

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COMPLETE SPECIFICATION

1 SHEET

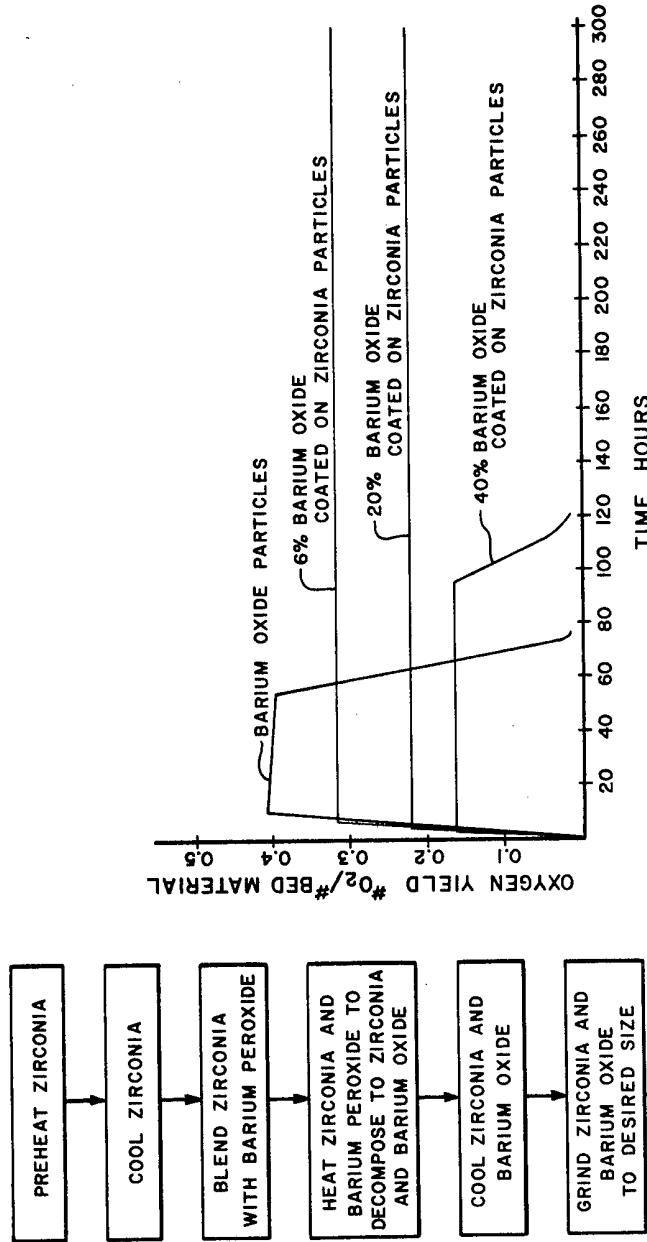
*This drawing is a reproduction of
the Original on a reduced scale*

FIG. 2

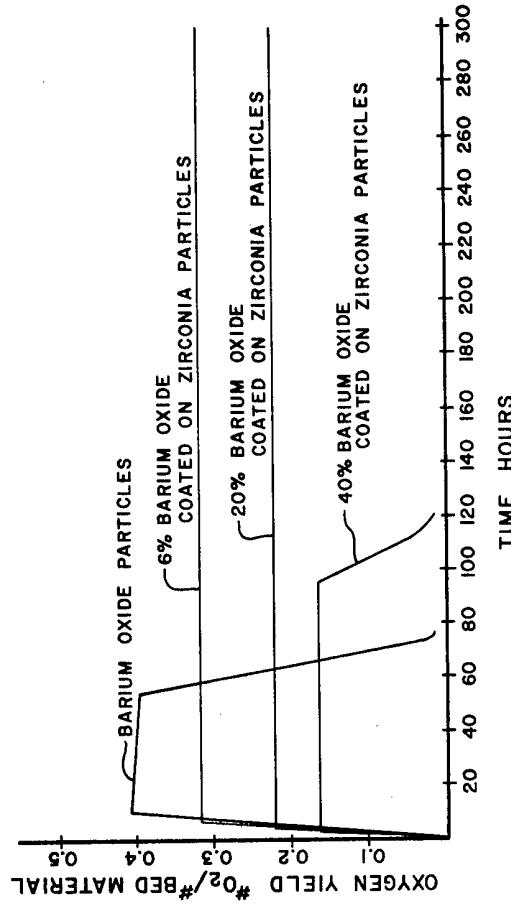


FIG. 1