**(12) STANDARD PATENT** (11) Application No. **AU 2020243365 B2**
**(19) AUSTRALIAN PATENT OFFICE**

(54) Title
**Methods of preparing fluorinated alcohols**

(51) International Patent Classification(s)
***C07D 301/26*** (2006.01) ***C07D 303/08*** (2006.01)

(21) Application No: **2020243365** (22) Date of Filing: **2020.03.18**

(87) WIPO No: **WO20/188274**

(30) Priority Data

| (31) Number | (32) Date | (33) Country |
|---|---|---|
| **1903909.8** | **2019.03.21** | **GB** |

(43) Publication Date: **2020.09.24**
(44) Accepted Journal Date: **2025.02.20**

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(56) Related Art
**Olah, G.A. and Meidar, D., 1978. Synthetic Methods and Reactions. 43. Preparation of Fluorohydrins from Epoxides with Pyridinium Poly#hydrogen Fluoride. Israel Journal of Chemistry, 17(1#2), 1978, pp.148-149.**
**BRYCE M R: "Reactions involving fluoride ion. Part 30. Preparation and reactions of epoxides derived from perfluoroalkyl substituted alkenes", JOURNAL OF THE CHEMICAL SOCIETY, PERKIN TRANSACTIONS 1, no. 7, 1 January 1984, pp 1391 - 1395,**
**KALOW J. A and DOYLE A. G: "Enantioselective Ring Opening of Epoxides by Fluoride Anion Promoted by a Cooperative Dual-Catalyst System", JOURNAL OF THE AMERICAM SOCIETY, vol. 132, no. 10, 17 March 2010 (2010-03-17), US, pages 3268 - 3269, X**
**WO 2018/197897 A1**

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization
International Bureau

(43) International Publication Date
24 September 2020 (24.09.2020)

WIPO | PCT

(10) International Publication Number
WO 2020/188274 A3

(51) International Patent Classification:
*C07D 301/26* (2006.01) *C07D 303/08* (2006.01)

(21) International Application Number:
PCT/GB2020/050701

(22) International Filing Date:
18 March 2020 (18.03.2020)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:
1903909.8 21 March 2019 (21.03.2019) GB

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(81) Designated States *(unless otherwise indicated, for every kind of national protection available)*: AE, AG, AL, AM, AO, AT, AU, AZ, BA, BB, BG, BH, BN, BR, BW, BY, BZ, CA, CH, CL, CN, CO, CR, CU, CZ, DE, DJ, DK, DM, DO, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, GT, HN, HR, HU, ID, IL, IN, IR, IS, JO, JP, KE, KG, KH, KN, KP, KR, KW, KZ, LA, LC, LK, LR, LS, LU, LY, MA, MD, ME, MG, MK, MN, MW, MX, MY, MZ, NA, NG, NI, NO, NZ, OM, PA, PE, PG, PH, PL, PT, QA, RO, RS, RU, RW, SA, SC, SD, SE, SG, SK, SL, ST, SV, SY, TH, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, WS, ZA, ZM, ZW.

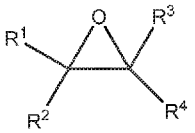
(84) Designated States *(unless otherwise indicated, for every kind of regional protection available)*: ARIPO (BW, GH, GM, KE, LR, LS, MW, MZ, NA, RW, SD, SL, ST, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, RU, TJ, TM), European (AL, AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HR, HU, IE, IS, IT, LT, LU, LV, MC, MK, MT, NL, NO, PL, PT, RO, RS, SE, SI, SK, SM, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, KM, ML, MR, NE, SN, TD, TG).

**Published:**
— *with international search report (Art. 21(3))*
— *before the expiration of the time limit for amending the claims and to be republished in the event of receipt of amendments (Rule 48.2(h))*

(88) Date of publication of the international search report:
12 November 2020 (12.11.2020)

WO 2020/188274 A3

(54) Title: METHODS OF PREPARING FLUORINATED ALCOHOLS




(57) Abstract: A method for preparing a partially fluorinated alcohol, comprises reacting an epoxide: wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl with a fluorinating agent.

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***Composition***

The present invention relates to methods of preparing partially fluorinated alcohols (fluorohydrins) from fluorinated epoxides and preparing fluorinated carbonate esters from fluorohydrins.

Fluorohydrins are useful as solvents and as synthetic building blocks from which various species such as esters, ethers, ketones, aldehydes and acids can be prepared. Of particular interest is their utility in the preparation of fluorinated carbonate esters, which are an important class of materials with significant commercial value. Fluorinated carbonate esters are commonly used without modification as synthetic intermediates and as solvents in electronic devices such as batteries (e.g. lithium ion batteries) and to manufacture products such as lubricants, sealants, and coatings.

The production of fluorohydrins from epoxides is known in the art. For example, Olah described a general method for preparing fluorohydrins by ring opening epoxides with a nucleophilic source of fluoride (G. A. Olah et al, Israel Jr. Chem., 17(1978),148-149). However, Olah did not extend this work to the preparation of fluorohydrins from fluorinated epoxides.

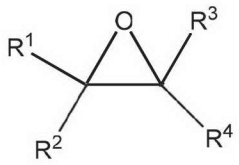
The ring opening of the fluorinated epoxide, 2,3-epoxy1,1,1-trifluoropropane (TFPO), with various nucleophiles to form fluorohydrins was generally described in a review of the chemistry of TFPO by Uneyama in Jr. Fluorine Chem., 105(2000) 285-293. However, this review was silent on the possibility or likely outcome of attempting to ring open TFPO or indeed any other fluorinated epoxide with nucleophilic fluorinating agents as taught by Olah.

General methods for the production of carbonate esters from alcohols and a carboxylating agent are known in the art, see for example "March's Advanced Organic Chemistry", M. B. Smith and J. March, 6$^{th}$ edition, page 1276. However, the production of fluorinated carbonate esters from fluorohydrins and carboxylating agents is unknown as are the products of such reactions.

*METHODS OF THE INVENTION*

According to a first aspect of the invention there is provided a method for preparing a partially fluorinated alcohol, comprising reacting a fluorinated epoxide:

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wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl with a fluorinating agent;

where one of of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -F and two of $R^1$ to $R^4$ are -H; or

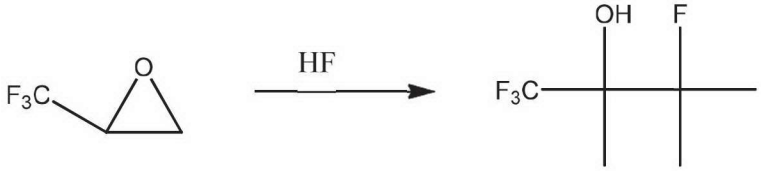
where two of $R^1$ to $R^4$ are $-CF_3$ and two of $R^1$ to $R^4$ are -H; or

where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -Cl and two of $R^1$ to $R^4$ are -H.

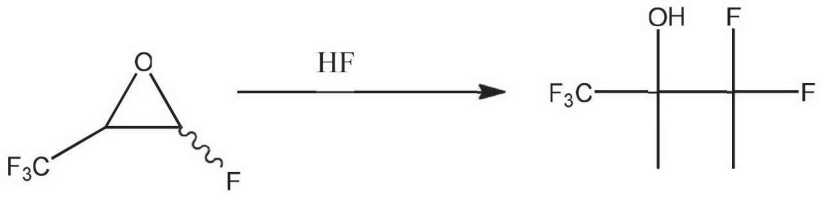
Preferably at least one of $R^1$ to $R^4$ comprises F, $CF_3$ or fluoroalkyl.

Preferably the fluorinating agent comprises a nucleophilic fluorinating agent. Preferred examples of fluorinating agents include HF and complexes of HF with nitrogen containing species, such as Olah's reagent (HF:Pyridine complex), with urea, or with a tertiary amine.

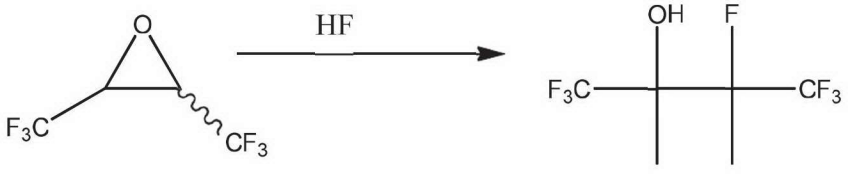
The method may comprise reacting an epoxide of 3,3,3-Trifluoropropene (1243zf) with HF and/or Olah's reagent to form $CF_3CH(OH)CH_2F$.



The method may comprise reacting an epoxide of 1,3,3,3-Tetrafluoropropene (1234ze) with HF and/or Olah's reagent to form $CF_3CH(OH)CHF_2$.

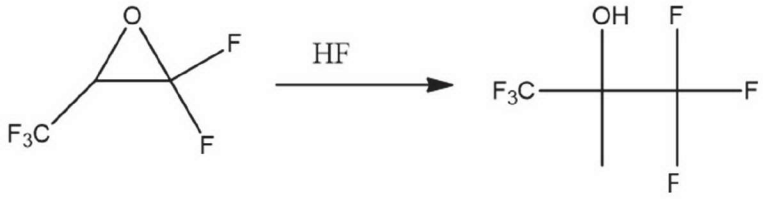


The method may comprise reacting an epoxide of 1,1,1,4,4,4-hexafluoro-2-butene (1336mzz) with HF and/or Olah's reagent to form $CF_3CH(OH)CHF(CF_3)$.



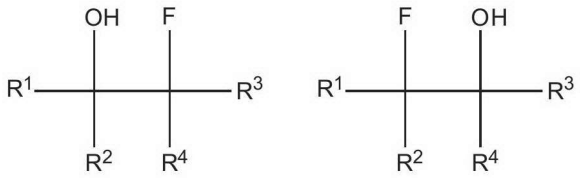
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The method may comprise reacting an epoxide of 1,1,3,3,3-Pentafluoropropene (1225zc) with HF and/or Olah's reagent to form $CF_3CH(OH)CF_3$.




*COMPOUNDS & COMPOSITIONS OF THE INVENTION*

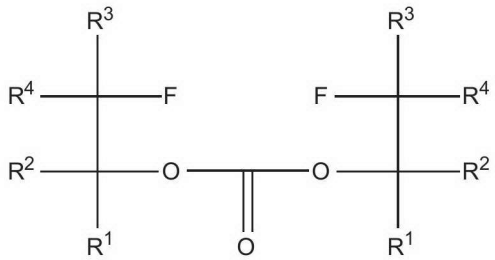
According to a second aspect of the invention there is provided compounds with the structure




wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl, with the provision that the compound is not 1,1,1,3-tetrafluoropropan-2-ol.

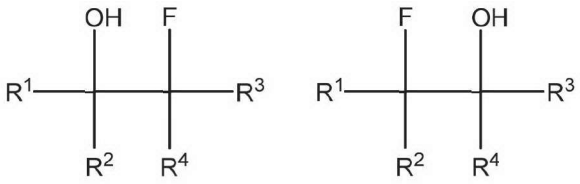
The compounds of the second aspect of the invention may be used in the preparation of a carbonate ester.

According to a third aspect of the invention there is provided a method for preparing a partially fluorinated carbonate ester with the structure




comprising reacting a fluorohydrin

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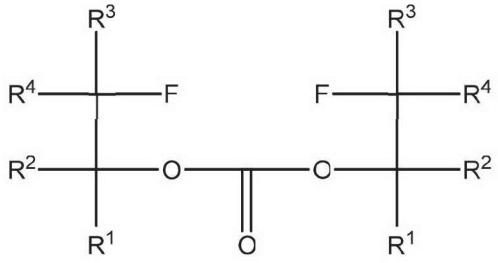



wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl with $COX_2$, wherein X is selected from the group comprising -F, -Cl, $-OCH_3$, $-OCCl_3$, imidazole, succinimidyl.

Preferably 2 equivalents (on a molar basis) of the fluorohydrin are used per 1 equivalent of $COX_2$.

Alternatively, 1 equivalent of the fluorohydrins of this invention maybe used with 1 equivalent of an alcohol species (a branched or linear monohydric/polyhydric alcohol) to prepare asymmetric carbonate esters.

The compounds produced in a method according to the third aspect of the invention are covered by the fourth aspect of the invention. According to the fourth aspect of the invention there is provided a compound with the structure




wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl.

The compounds of the fourth aspect of the invention may be also used as a battery solvent component (e.g. in a lithium ion battery). Here the compounds are found to be beneficial as a result of their physical properties, electrochemical stability, compatibility with battery components such as battery electrodes (cathodes and anodes) including electrodes comprising carbon and silicon, lithium containing electrolyte salts, separators, binders, current collectors and low flammability.

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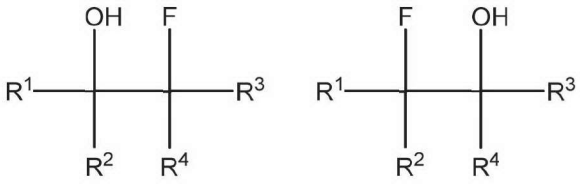
The compounds of the fourth aspect of the invention may also be used with other solvents and additives such as other linear and cyclic carbonate esters.

Preferably when used as a solvent the composition comprises an electrolyte salt. Preferred examples of electrolyte salts include lithium-based electrolytes such as those selected from the group comprising lithium hexafluorophosphate ($LiPF_6$), lithium triflate ($LiSO_3CF_3$), lithium bis(fluorosulfonyl)imide ($Li(FSO_2)_2N$) and lithium bis(trifluoromethanesulfonyl)imide ($Li(CF_3SO_2)_2N$).

The compounds of the second aspect of the invention may be used in the preparation of a (more highly) fluorinated derivative. One or more of the R groups may be substituted by fluorine. In the process the R groups to be altered by fluorination are preferably selected from the group comprising H, Cl, Br, I.

Depending on the nature of the R group(s) being modified the preparation process for the fluorinated derivative may comprise a multi-stage process; preferably a two-stage process. In a preferred two stage process a first stage is the modification of the targeted R group(s) to a (different) halogen group, preferably to a chlorine group (with a suitable chlorinating agent such as chlorine); in a second stage the chlorine group is modified to a fluorine group (with a suitable fluorinating agent such as HF or a metal fluorine salt, such as NaF, KF). It will be appreciated that where the R group(s) targeted already comprises a halogen, other than fluorine, a two-stage process with substitution of the halogen with chlorine may not be necessary.

Thus compounds



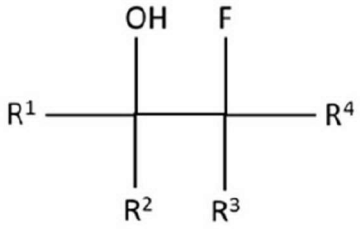

wherein at least 2 of $R^1$ to $R^4$ independently comprises H, Cl, Br, I may be converted to (more highly) fluorinated derivatives.

In the fluorinated derivatives preferably at least 2 and more preferably at least 3 of $R^1$ to $R^4$ independently comprises F, $CF_3$ or a fluoroalkyl. Preferably at least 1 of $R^1$ to $R^4$, and more preferably 1 of $R^1$ to $R^4$ independently comprises H. Most preferably 1 of $R^1$ to $R^4$ comprises

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$CF_3$, two of $R^1$ to $R^4$ comprise F and one of $R^1$ to $R^4$ comprises H. Most preferably the fluorinated derivative comprises hexfluroroisopropanol.

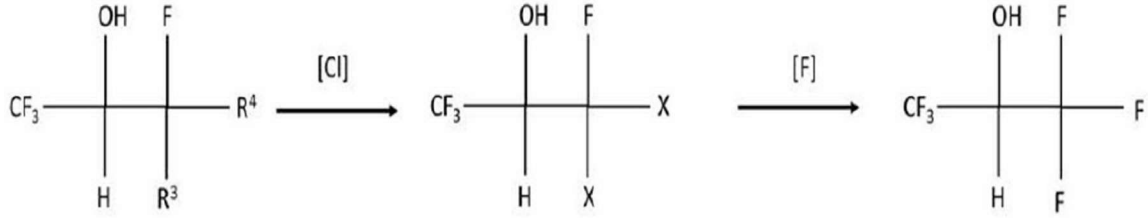
A preferred reaction pathway occurs for the compound




wherein in the preferred alternatives

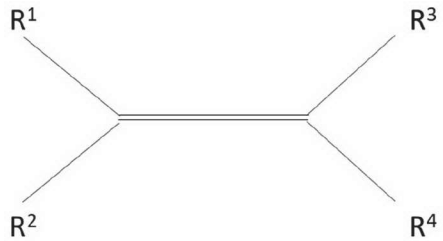
- $R^1$ is $-CF_3$, $R^2$ is H and both $R^3$ and $R^4$ are H.
- $R^1$ is $-CF_3$, $R^2$ is H; one of $R^3$ and $R^4$ is H, one of $R^3$ and $R^4$ is F.

This preferred pathway is shown below.




X is either F or Cl.

The epoxides useful in the first aspect of the invention may prepared from a fluorinated alkene. According to a fifth aspect of the invention there is provided a method for preparing a partially fluorinated epoxide, comprising reacting a fluorinated alkene:




wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl with an oxidising agent.

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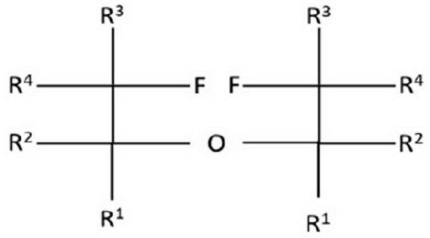
Preferably at least one of $R^1$ to $R^4$ comprises F, $CF_3$ or fluoroalkyl.

Preferred examples of oxidising agent include air, oxygen and oxygen containing compounds such as peroxides, per-salts and compounds of oxygen with other elements such as hypohalites. Preferably the oxidising agent comprises a hypohalite such as chlorite.

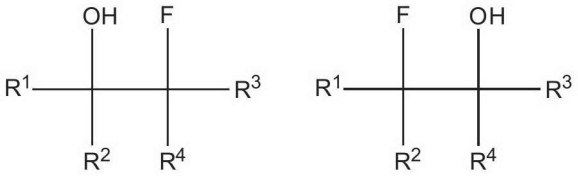
Preferably, the compound reacted with the oxidising agent is a tetrafluoropropene. Most preferably, one of $R^1$ and $R^2$ is $-CF_3$ and one of $R^3$ and $R^4$ is -F. Thus, the tetrafluoropropene is 1,3,3,3-Tetrafluoropropene (1234 ze) or 2,3,3,3-Tetrafluoropropene (1234yf).

According to a sixth aspect of the invention there is provided a method for preparing a fluorohydrin comprising the fifth and the first aspects of the invention.

According to a seventh aspect of the invention there is provided a method for preparing a partially fluorinated ether with the structure

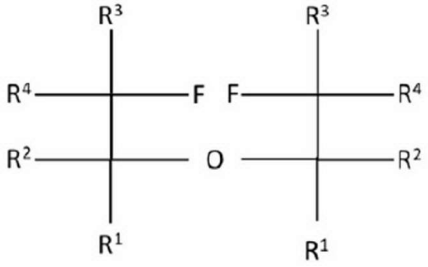



comprising reacting a fluorohydrin with the structure




wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl.

According to an eighth aspect of the invention there is provided a compound with the structure



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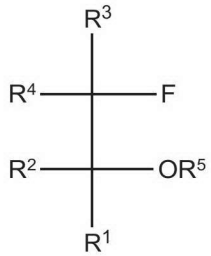
wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl.

According to a ninth aspect of the invention there is provided a composition comprising a compound of the eighth aspect of the invention.

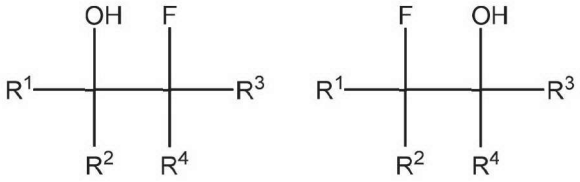
The compound of the eighth aspect of the invention or the composition according to the ninth aspect of the invention may be used as a solvent, for example, in battery applications.

The compound of the eight aspect of the invention or the composition according to the ninth aspect of the invention may be used as a coolant, for example, as an immersive coolant.

Also provided is a method for preparing a partially fluorinated ether with the structure




comprising reacting a fluorohydrin with the structure




wherein $R^1$, $R^2$, $R^3$, $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl and $R^5$ is independently selected from the group $CF_3$, alkyl, fluoroalkyl, perfluoroalkyl, haloalkyl perfluorohaloalkyl.

Preferably the ether synthesis occurs *via* acid catalysed dehydration of the fluorohydrin.

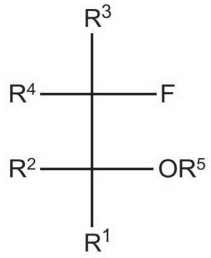
Alternatively the ether synthesis occurs *via* one or more of the following techniques:

i. Alkoxy-de-halogenation - reaction of an alkyl halide with a fluorohydrin, preferably under basic conditions;

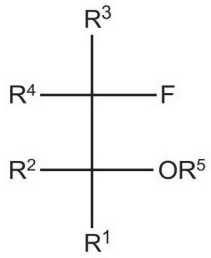
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ii. Alkoxy-de-sulphonyloxy-substitution - reaction of an fluorohydrin sulphate with an alkoxide or fluorohydrin alkoxide;
iii. Hydro, alkoxy-de-diazo-disubstitution - reaction of a fluorohydrin with a diazo compound;
iv. Alkoxy-de-hydroxylation - dehydration of two alcohols to yield an ether with e.g. concentrated sulphuric acid;
v. Hydroxy or alkoxy-de-alkoxylation - transetherification of a (fluorohydrin)ether with an alcohol or fluorohydrin; and / or
vi. Alkoxy-de-hydroxylation - reaction of an alcohol or fluorohydrin with an oxonium compound.

Additionally provided is a compound with the structure




Further provided is a composition comprising a compound with the structure




The compound or the composition may be used as a solvent, for example, in battery applications.

The compound or the composition may be used as a coolant, for example, as an immersive coolant.

The invention will now be illustrated with reference to the following non-limiting examples.

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***Examples***

***Example 1 – Ring opening of an epoxide with Olah's reagent***

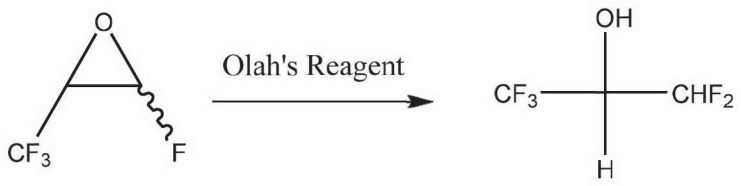
The following steps were followed.

- The reactor was charged with Olah's reagent (70 % HF:Pyridine, 5 ml) and cooled in an ice batch with stirring.
- 2,3-epoxy1,1,1-trifluoropropane (TFPO) (3.4 g) was then added dropwise.
- At the end of the addition the reaction mixture was allowed to warm up to room temperature; stirring was continued for 48 hours.
- After 48 hours the reaction mixture was quenched with ice.
- Salt was added, and the product extracted with diethyl ether (3 x 5 ml). The diethyl ether extracts were combined, washed with saturated potassium bicarbonate solution and water before being dried over anhydrous sodium sulphate. Diethyl ether was removed in vacuo to yield the desired product as a clear, colourless liquid boiling point 91-93$^{O}$C. The identity of this product was confirmed by NMR spectroscopy.

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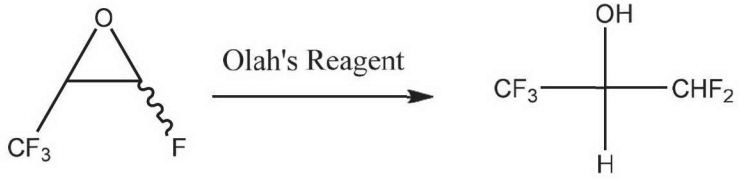
***Example 2 – Ring opening of* 2,3-epoxy-1,1,1,3-tetrafluoropropane *with Olah's reagent***




2,3-epoxy-1,1,1,3-tetrafluoropropane was ring opened using the following procedure:

- A 100 ml Hastalloy C pressure reactor was charged with Olah's reagent (70 % HF:Pyridine, 25 g).
- After sealing, the contents of the reactor were cooled to 20$^{O}$C with stirring.
- 2,3-epoxy-1,1,1,3-tetrafluoropropane (11 g) was then added.
- After this addition was complete the reaction mixture was heated to 50$^{O}$C and stirred for 168 hours.
- After 168 hours the reaction mixture was quenched with ice and saturated sodium chloride solution (22 ml) added.
- The product was extracted from this mixture with diethyl ether.
- The diethyl ether extracts were combined, washed with saturated potassium bicarbonate solution and then water before being dried over anhydrous sodium sulphate. The identity of the product was confirmed by NMR spectroscopy.

***Example 2a – Ring opening of* 2,3-epoxy-1,1,1,3-tetrafluoropropane *with Olah's reagent***

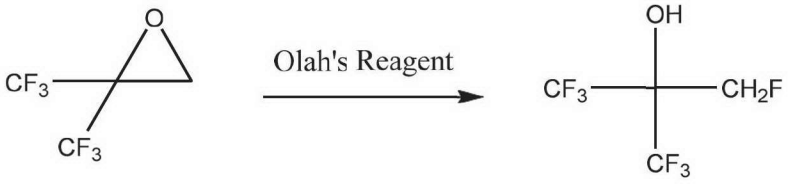



2,3-epoxy-1,1,1,3-tetrafluoropropane was ring opened using the following procedure:

- A 100 ml Hastalloy C pressure reactor was charged with Olah's reagent (70 % HF:Pyridine, 25 g).
- After sealing, the contents of the reactor were cooled to 20$^{O}$C with stirring.
- 2,3-epoxy-1,1,1,3-tetrafluoropropane (10.6 g) was then added.
- After this addition was complete the reaction mixture was heated to 80$^{O}$C and stirred for 43 hours.
- After 43 hours a sample of the reaction mixture was analysed by GCMS and it was found that all the feed epoxide had reacted.
- After cooling the reaction mixture was quenched with ice and saturated sodium chloride solution (22 ml) added.
- The product was extracted from this mixture with diethyl ether.
- The diethyl ether extracts were combined, washed with saturated potassium bicarbonate solution and then water before being dried over anhydrous sodium sulphate. The identity of the product was confirmed by NMR spectroscopy.

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***Example 3 – Ring opening of* 2,3-epoxy-1,1,1-trifluoro-2-(trifluoromethyl)propane *with Olah's reagent***




2,3-epoxy-1,1,1-trifluoro-2-(trifluoromethyl)propane was ring opened using the following procedure:

- A 100 ml Hastalloy C pressure reactor was charged with Olah's reagent (70 % HF:Pyridine, 16.5 g).
- After sealing, the contents of the reactor were cooled to 20$^{O}$C with stirring.
- 2,3-epoxy-1,1,1-trifluoro-2-(trifluoromethyl)propane (10 g) was then added.
- After this addition was complete the reaction mixture was heated to 50$^{O}$C and stirred for 160 hours.
- After 160 hours the reaction mixture was quenched with ice and saturated sodium chloride solution (22 ml) added.
- The product was extracted from this mixture with diethyl ether.
- The diethyl ether extracts were combined, washed with saturated potassium bicarbonate solution and then water before being dried over anhydrous sodium sulphate. The identity of the product was confirmed by NMR spectroscopy.

***Example 4 – Preparation of* di-(1,1,1,3-tetrafluoropropyl) carbonate *with phosgene***

Di-(1,1,1,3-tetrafluoropropyl) carbonate was synthesised using the following procedure:

- A three necked round bottom flask was cooled to 0$^{O}$C under an inert atmosphere.
- Phosgene solution (15% by weight in toluene, 50 mL of solution) was added and stirred.
- A mixture of 1,1,1,3-tetrafluoropropan-2-ol (18.42 g) and pyridine (11.02 g) was added to the solution dropwise, and the temperature of the solution was monitored to ensure it did not rise above 10$^{O}$C.
- The solution was allowed to warm to room temperature and stirred for 48 hours.
- The product was filtered to remove pyridinium salts and the solvent was removed in vacuo to yield the crude product.
- The crude product was distilled under atmospheric conditions to yield di-(1,1,1,3-tetrafluoropropyl) carbonate as a yellow oil (7.08 g, 35 % yield).

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***Figures***

The Figures illustrates the results of various spectroscopic analytical techniques carried out on some of the reaction products from the Examples.

Figure 1 shows a $^{19}F$ NMR spectrum of the reaction product of 2,3-epoxy1,1,1-trifluoropropane (TFPO) with Olah's reagent.

Figure 2a shows a $^{19}F$ NMR spectrum of the reaction product of 2,3-epoxy-1,1,1,3-tetrafluoropropane ring opening with Olah's reagent.

Figure 2b shows a proton coupled $^{19}F$ NMR spectrum of the reaction product of 2,3-epoxy-1,1,1,3-tetrafluoropropane ring opening with Olah's reagent.

Figure 2c shows a proton decoupled $^{19}F$ NMR spectrum of the reaction product of 2,3-epoxy-1,1,1,3-tetrafluoropropane ring opening with Olah's reagent.

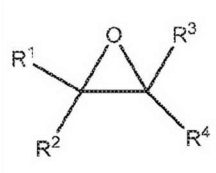
Figure 3 shows a $^{19}F$ NMR spectrum of the reaction product of 2,3-epoxy-1,1,1-trifluoro-2-(trifluoromethyl)propane) ring opening with Olah's reagent.

Figure 4 shows a $^{19}F$ NMR spectrum of the reaction product of 1,1,1,3-tetrafluoropropan-2-ol with phosgene, consistent with that of the product di-(1,1,1,3-tetrafluoropropyl) carbonate.

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Claims:

1. A method for preparing a partially fluorinated alcohol, comprising reacting an epoxide:




wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl with a fluorinating agent;

where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -F and two of $R^1$ to $R^4$ are -H; or

where two of $R^1$ to $R^4$ are $-CF_3$ and two of $R^1$ to $R^4$ are -H; or

where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -Cl and two of $R^1$ to $R^4$ are -H.

2. A method according to claim 1, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -F and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is -F and $R^3$ is $-CF_3$, $R^2$ and $R^4$ are -H.

3. A method according to claim 1, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -F and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is -F and $R^2$ is $-CF_3$, $R^3$ and $R^4$ are -H.

4. A method according to claim 1, where two of $R^1$ to $R^4$ are $-CF_3$ and two of $R^1$ to $R^4$ are -H , and wherein $R^1$ is $-CF_3$ and $R^3$ is $-CF_3$; $R^2$ is -H and $R^4$ are -H.

5. A method according to claim 1, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -Cl and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is -Cl and $R^3$ is $-CF_3$, $R^2$ and $R^4$ are – H.

6. A method according to claim 1, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -Cl and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is -Cl and $R^2$ is $-CF_3$, $R^3$ and $R^4$ are -H.

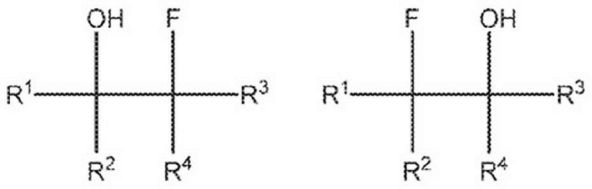
7. A method according to any of the preceding claims when dependent on claim 1, where the fluorinating agent comprises a nucleophilic fluorinating agent.

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8. A method according to claim 1 or 7, where the fluorinating agent is selected from HF and complexes of HF with nitrogen containing species, such as Olah's reagent (HF: Pyridine complex), with urea, or with a tertiary amine.

9. A method according to claim 8, where the HF: Pyridine complex has a ratio of 7:3 by weight.

10. Compounds with the structure




wherein $R^1$, $R^2$, $R^3$ and $R^4$ are independently selected from the group comprising H, F, Cl, Br, I, $CF_3$, alkyl, fluoroalkyl, haloalkyl;
where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -F and two of $R^1$ to $R^4$ are -H; or
where two of $R^1$ to $R^4$ are $-CF_3$ and two of $R^1$ to $R^4$ are -H; or
where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -Cl and two of $R^1$ to $R^4$ are -H;
formed by the method of any one of the preceding claims.

11. A compound according to claim 10, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -F and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is -F and $R^3$ is $-CF_3$, $R^2$ and $R^4$ are -H.

12. A compound according to claim 10, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -F and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is -F and $R^2$ is $-CF_3$, $R^3$ and $R^4$ are -H.

13. A compound according to claim 10, where two of $R^1$ to $R^4$ are $-CF_3$ and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is $-CF_3$ and $R^3$ is $-CF_3$; $R^2$ is -H and $R^4$ are -H.

14. A compound according to claim 10, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -Cl and two of $R^1$ to $R^4$ are -H, and wherein $R^1$ is -Cl and $R^3$ is $-CF_3$, $R^2$ and $R^4$ are – H.

15. A compound according to claim 10, where one of $R^1$ to $R^4$ is $-CF_3$, one of $R^1$ to $R^4$ is -Cl and two of $R^1$ to $R^4$ are -H, and wherein preferably $R^1$ is -Cl and $R^2$ is $-CF_3$, $R^3$ and $R^4$ are -H.

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16. Use of a compound according to claim 10 as a solvent, in battery applications.

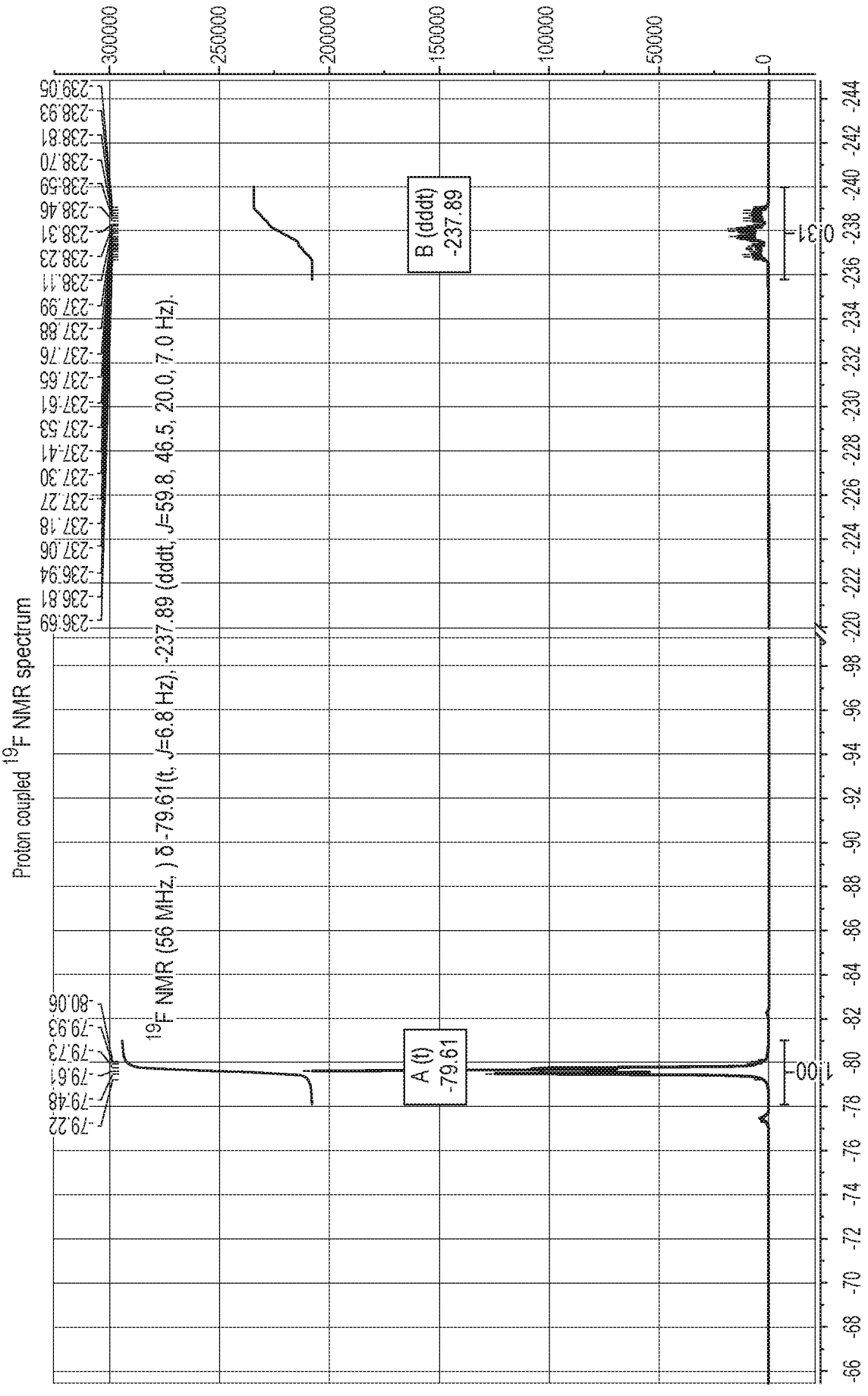

Proton coupled $^{19}$F NMR spectrum
$^{19}$F NMR (56 MHz, ) δ -79.61(t, J=6.8 Hz), -237.89 (dddt, J=59.8, 46.5, 20.0, 7.0 Hz).
A (t)
-79.61
B (dddt)
-237.89
f1 (ppm)

FIG. 1

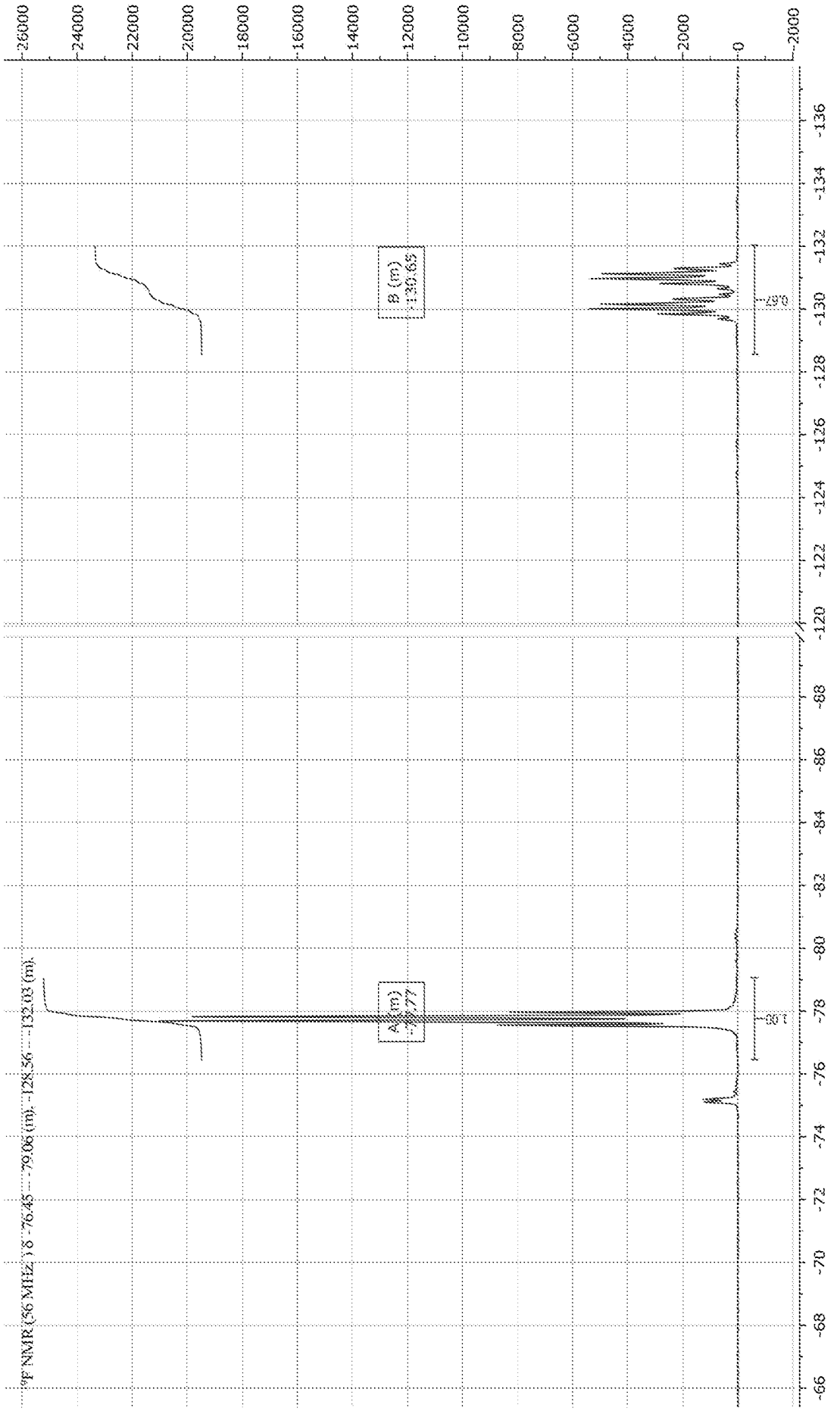



Figure 2: Proton coupled $^{19}F$ NMR spectrum

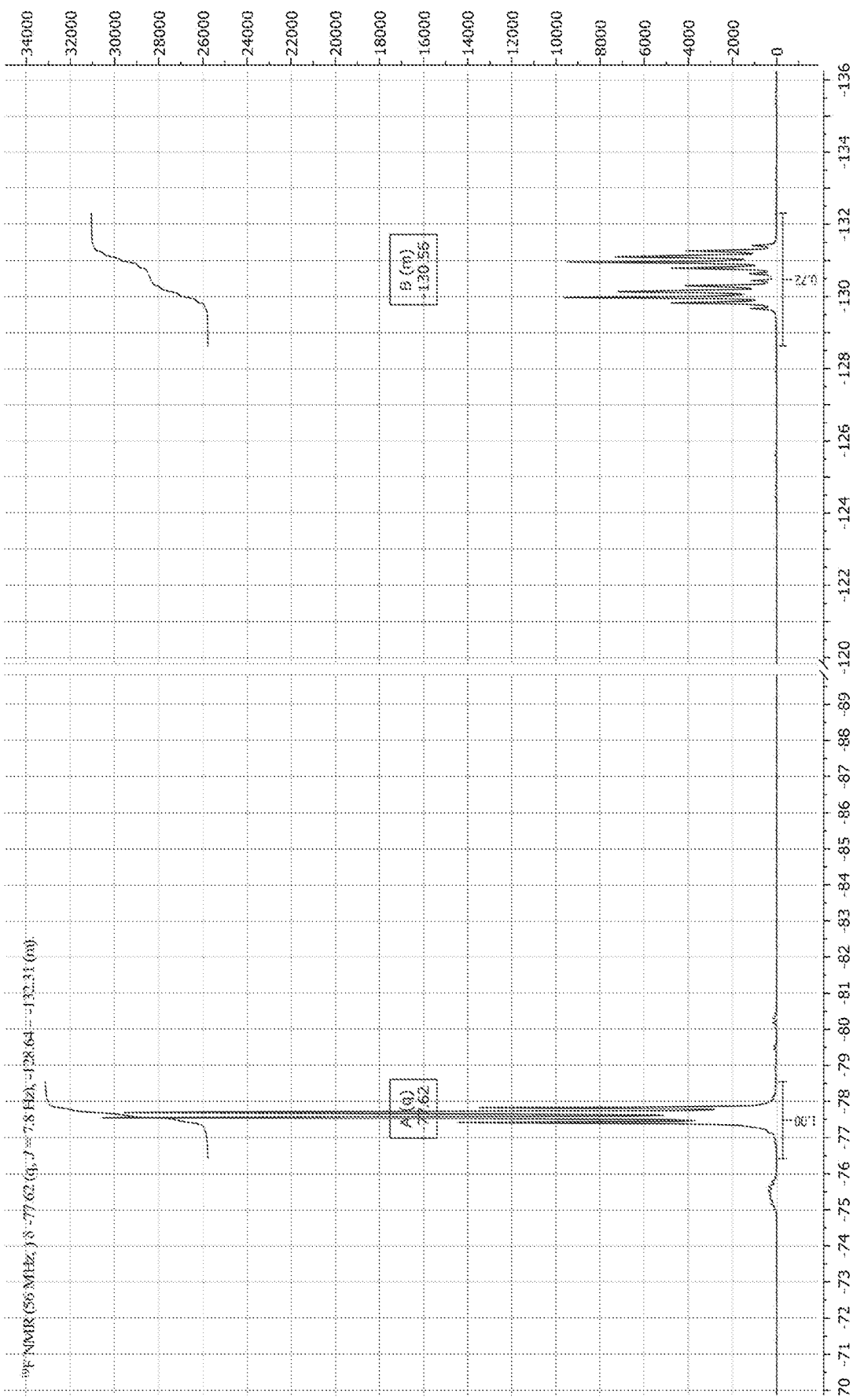



Figure 2a: Proton coupled $^{19}F$ NMR spectrum

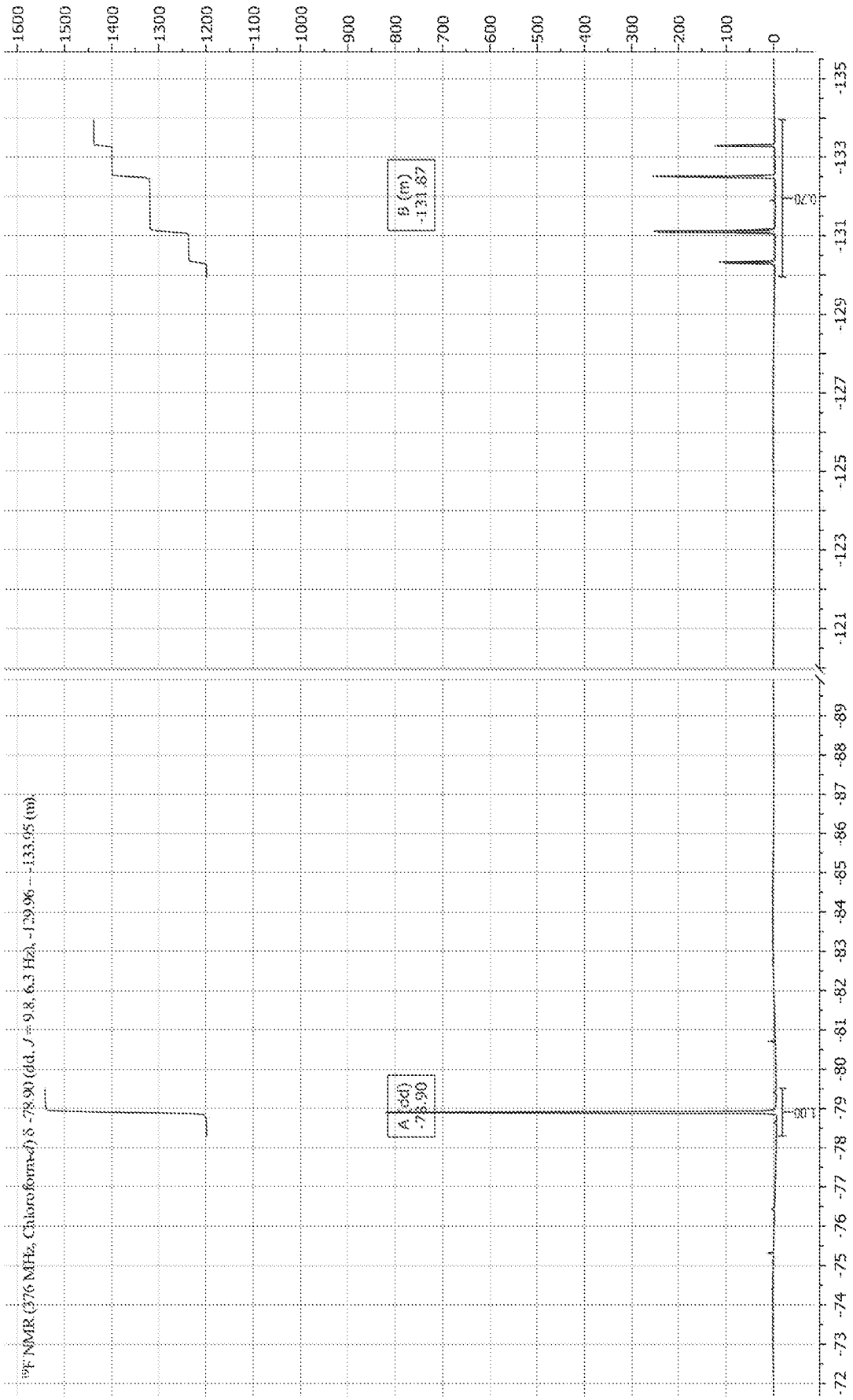



Figure 2a: Proton decoupled $^{19}$F NMR spectrum

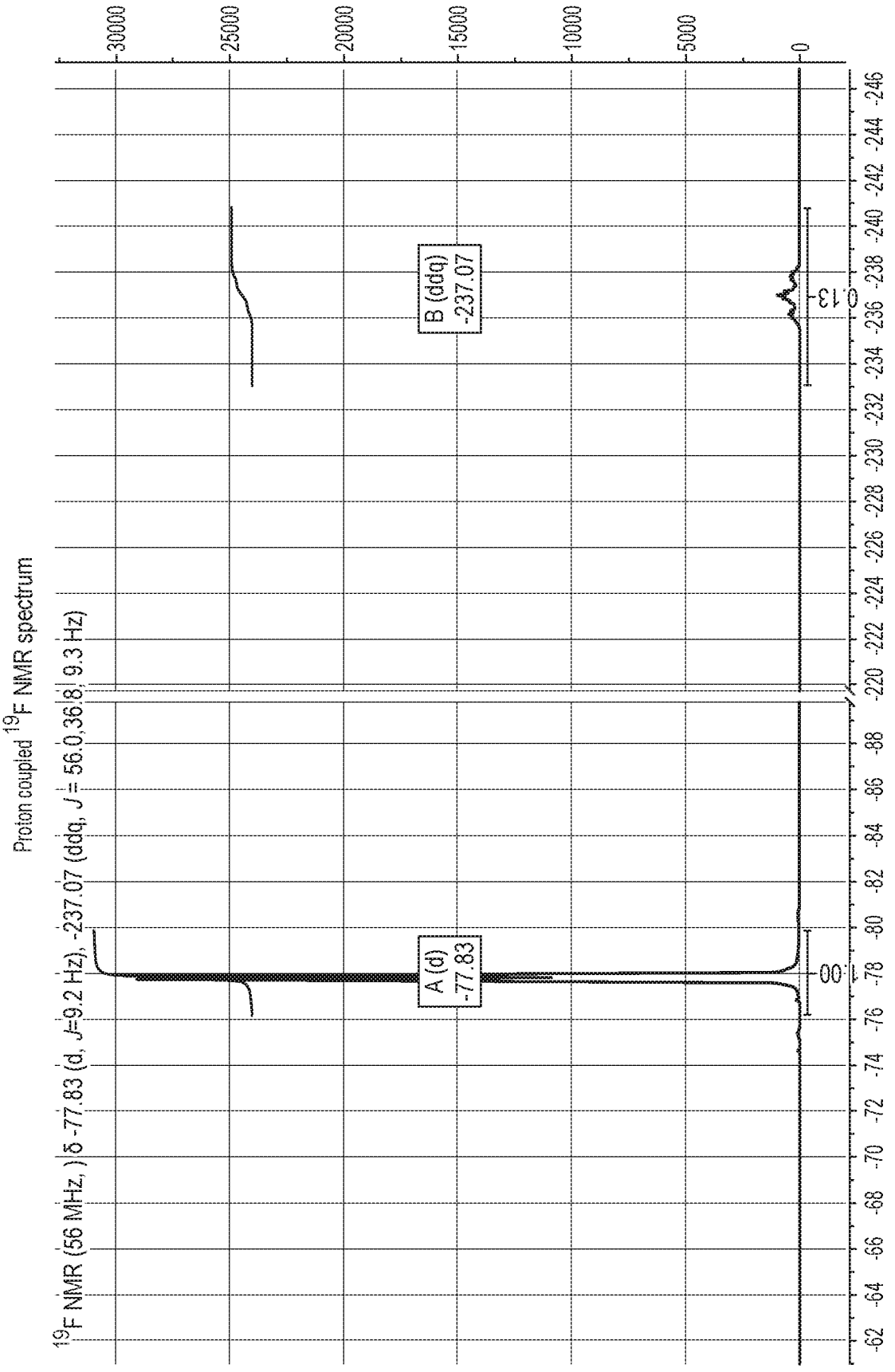

Proton coupled $^{19}$F NMR spectrum
$^{19}$F NMR (56 MHz, ) δ -77.83 (d, $J$=9.2 Hz), -237.07 (ddq, $J$ = 56.0,36.8, 9.3 Hz)
A (d)
-77.83
B (ddq)
-237.07
30000
25000
20000
15000
10000
5000
0
-62
-64
-66
-68
-70
-72
-74
-76
-78
-80
-82
-84
-86
-88
-220
-222
-224
-226
-228
-230
-232
-234
-236
-238
-240
-242
-244
-246

FIG. 3

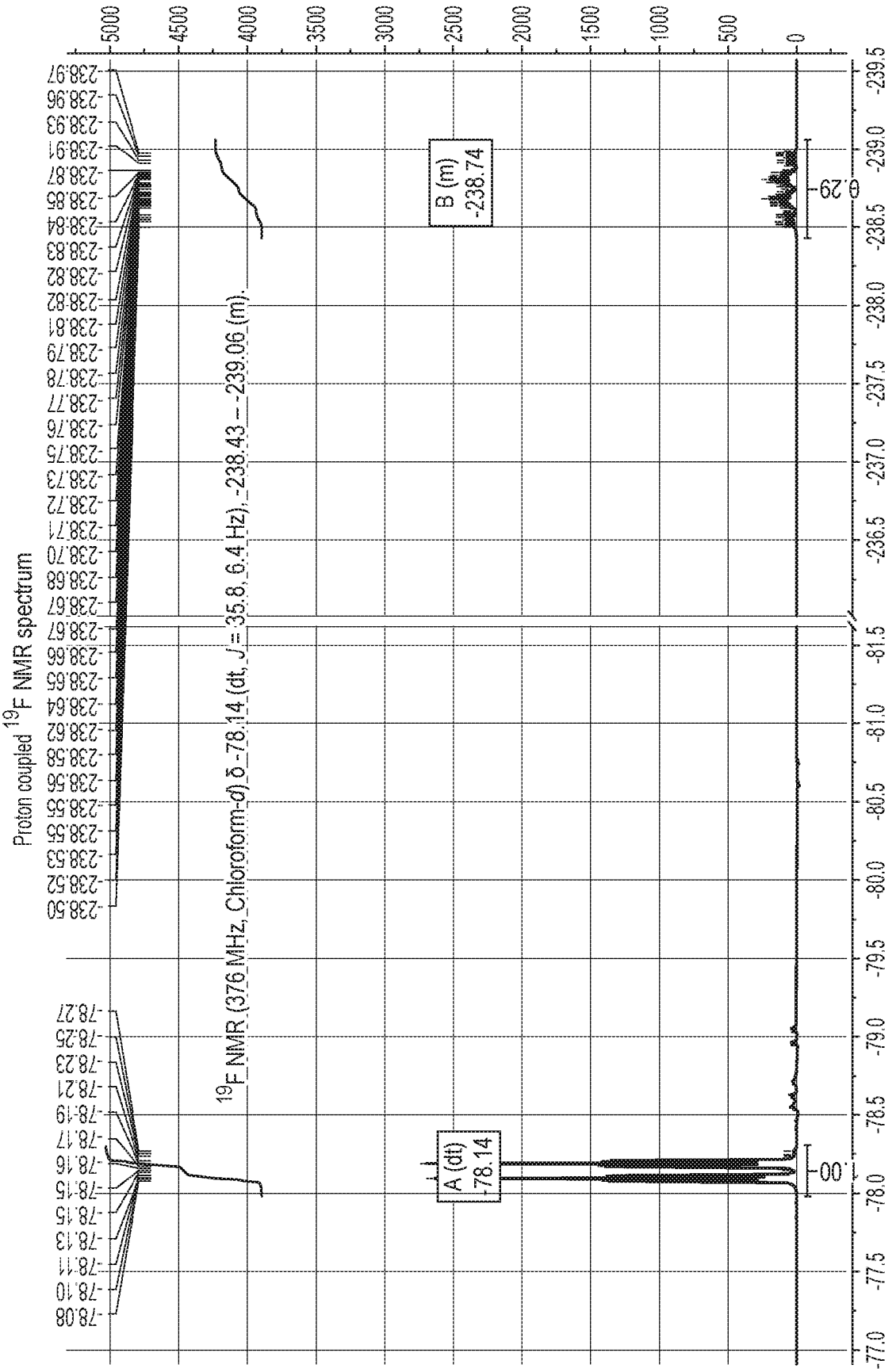

Proton coupled $^{19}$F NMR spectrum
$^{19}$F NMR (376 MHz, Chloroform-d) δ -78.14 (dt, J = 35.8, 6.4 Hz), -238.43 – -239.06 (m).
A (dt) -78.14
B (m) -238.74
1.00
0.29

FIG. 4