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(54) Title: NANOWIRE PREPARATION METHODS, COMPOSITIONS, AND ARTICLES

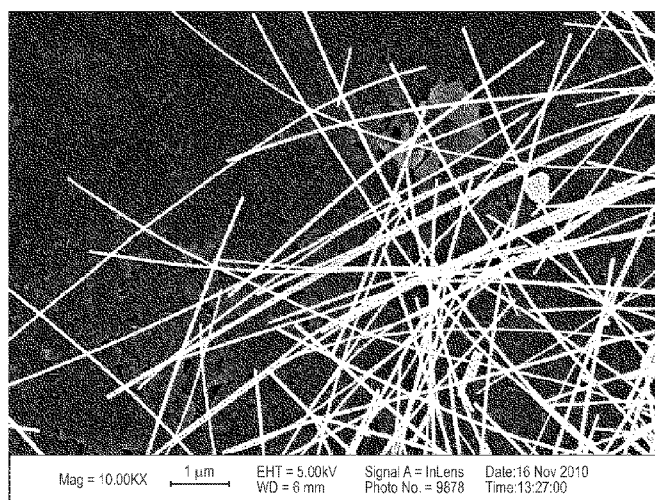


FIG. 2

(57) Abstract: Nanowire preparation methods, compositions, and articles are disclosed. Such methods which reduce metal ions to metal nanowires in the presence of aluminum or gallium ions, are capable of producing long, narrow, nanowires useful for electronics and optical applications.

NANOWIRE PREPARATION METHODS,
COMPOSITIONS, AND ARTICLES

BACKGROUND

5 The general preparation of silver nanowires (10-200 aspect ratio) is known. See, for example, *Angew. Chem. Int. Ed.* 2009, 48, 60, Y. Xia, Y. Xiong, B. Lim, S. E. Skrabalak, which is hereby incorporated by reference in its entirety. Such preparations typically employ Fe^{2+} or Cu^{2+} ions to “catalyze” the wire formation over other morphologies. The controlled preparation of silver
10 nanowires having the desired lengths and widths, however, is not known. For example, the Fe^{2+} produces a wide variety of lengths or thicknesses and the Cu^{2+} produces wires that are too thick for many applications.

 When iron or copper are used, they are typically provided as the metal halide salts FeCl_2 or CuCl_2 . See, for example, B. Wiley et al., *Nano Letters*,
15 **2004**, 4, 1733-1739 and K.E. Korte et al., *J. Mats. Chem.*, **2008**, 18, 437. Other metal halide salts have been used in nanowire synthesis. See, for example, J. Jiu, K. Murai, D. Kim, K. Kim, K. Suganuma, *Mat. Chem. & Phys.*, **2009**, 114, 333, which refers to NaCl , CoCl_2 , CuCl_2 , NiCl_2 and ZnCl_2 , and S. Nandikonda, “Microwave Assisted Synthesis of Silver Nanorods,” M.S. Thesis, Auburn
20 University, Auburn, Alabama, USA, August 9, 2010, which refers to NaCl , KCl , MgCl_2 , CaCl_2 , MnCl_2 , CuCl_2 , and FeCl_3 , and Japanese patent application publication 2009-155674, which discloses SnCl_4 . See also S. Murali et al., *Langmuir*, **2010**, 26(13), 11176-83; Z. C. Li et al., *Micro & Nano Letters*, **2011**, 6(2), 90-93; and B. J. Wiley et al., *Langmuir*, **2005**, 21, 8077. Japanese patent
25 application publication 2009-155674 discloses SnCl_4 .

SUMMARY

 At least a first embodiment provides methods comprising providing at least one second metal or metal ion comprising at least one of an aluminum ion
30 or a gallium ion; and reducing at least one first reducible metal ion to at least one first metal nanowire, where the at least one second metal or metal ion and the at least one first reducible metal ion differ in atomic number. In at least some such

embodiments, the reducing the at least one first reducible metal ion occurs in the presence of the at least one second metal or metal ion.

In some cases, the providing the at least one second metal or metal ion comprises providing a composition comprising at least one first compound
5 comprising the at least one first reducible metal ion; at least one second compound comprising the at least one second metal or metal ion; and at least one solvent. In such methods, the at least one second compound may, for example, comprise at least one chloride ion. Gallium(III) chloride and aluminum(III) chloride hexahydrate are exemplary second compounds.

10 In at least some embodiments, the at least one first reducible metal ion comprises at least of a coinage metal ion or an ion of an element from IUPAC Group 11 or at least one ion of silver.

In at least some embodiments, the at least one second metal or metal ion comprises gallium or an ion of gallium. In other embodiments, the at
15 least one second metal or metal ion comprises aluminum or an ion of aluminum. In some cases, the at least one second metal or metal ion comprises gallium or an ion of gallium or aluminum or an ion of aluminum.

Some embodiments provide the first metal nanowire produced according to such methods.

20 At least a second embodiment provide a method comprising providing a composition comprising at least one first compound comprising at least one reducible metal ion, at least one second compound comprising at least one second metal or metal ion differing in atomic number from the at least one first reducible metal ion, the second metal or metal ion comprising at least one
25 element from IUPAC Group 13, and at least one solvent; and reducing the at least one first reducible metal ion to at least one first metal. The at least one first reducible metal ion may, for example, comprise at least one coinage metal ion, or at least one ion of an element from IUPAC Group 11, or at least one silver ion. The at least one first compound may, for example, comprise silver nitrate. The at
30 least one second metal or metal ion may, for example, comprise aluminum or an ion of aluminum or gallium or an ion of gallium. The at least one second compound may, for example, comprise at least one chloride, such as, for example,

aluminum(III) chloride hexahydrate or gallium(III) chloride. The at least one solvent may, for example, comprise at least one polyol, or at least one of ethylene glycol, propylene glycol, glycerol, one or more sugars, or one or more carbohydrates. In some embodiments, the ratio of the total moles of the at least one second metal or metal ion to the total moles of the at least one first reducible metal ion is from about 0.0001 to about 0.1. The reduction may, for example, be carried out at one or more temperatures from about 120 °C to about 190 °C, or from about 90 °C to about 190 °C.

In at least some embodiments, the composition further comprises at least one protecting agent. In some cases, the at least one protecting agent comprises at least one of: one or more surfactants, one or more acids, or one or more polar polymers. The at least one protecting agent may, for example, comprise polyvinylpyrrolidinone. At least some embodiments further comprise inerting the at least one protecting agent.

Other embodiments further comprise inerting one or more of the composition, the at least one compound comprising the at least one first reducible metal ion, the at least one second metal or metal ion, or the at least one solvent.

Still other embodiments provide the at least one first metal produced by such methods and articles comprising such at least one first metal. Such at least one first metal may, for example, comprise one or more nanowires, nanocubes, nanorods, nanopyramids, nanotubes, and the like. Such at least one first metal may comprise at least one object having an average diameter of between 10 nm and about 500 nm, or having an aspect ratio from about 50 to about 10,000.

Yet still other embodiments provide at least one metal nanowire with an average diameter between about 10 nm and about 150 nm, with an aspect ratio from about 50 to about 10,000. Such nanowires may, for example, comprise at least one coinage metal, or at least one element of IUPAC Group 11, or silver. Yet another embodiment comprises at least one article comprising such metal nanowires, such as, for example, at least one electronic device.

These embodiments and other variations and modifications may be better understood from the brief description of figures, description, exemplary

embodiments, examples, figures, and claims that follow. Other desirable objectives and advantages inherently achieved may occur or become apparent to those skilled in the art. The invention is defined by the appended claims.

5 **BRIEF DESCRIPTION OF FIGURES**

Figure 1 shows an optical micrograph of the silver nanowire product of Example 1.

Figure 2 shows a scanning electron micrograph of the purified silver nanowire product of Example 1.

10 Figure 3 shows an optical micrograph of the silver nanowire product of Example 2.

DESCRIPTION

15 All publications, patents, and patent documents referred to in this document are incorporated by reference herein in their entirety, as though individually incorporated by reference.

U.S. provisional application No. 61/488,821, filed May 23, 2011, entitled NANOWIRE PREPARATION METHODS, COMPOSITIONS, AND ARTICLES, is hereby incorporated by reference in its entirety.

20

Reducible Metal Ions and Metal Products

Some embodiments provide methods comprising reducing at least one reducible metal ion to at least one metal. A reducible metal ion is a cation that is capable of being reduced to a metal under some set of reaction conditions.

25 In such methods, the at least one first reducible metal ion may, for example, comprise at least one coinage metal ion. A coinage metal ion is an ion of one of the coinage metals, which include copper, silver, and gold. Or such a reducible metal ion may, for example, comprise at least one ion of an IUPAC Group 11 element. IUPAC Group 11 elements are sometimes referred to as Group IB elements, based on historic nomenclature. An exemplary reducible metal ion is a
30 silver cation. Such reducible metal ions may, in some cases, be provided as salts. For example, silver cations might, for example, be provided as silver nitrate.

In such embodiments, the at least one metal is that metal to which the at least one reducible metal ion is capable of being reduced. For example, silver would be the metal to which a silver cation would be capable of being reduced.

5

Nanostructures, Nanostructures, and Nanowires

In some embodiments, the metal product formed by such methods is a nanostructure, such as, for example, a one-dimensional nanostructure.

Nanostructures are structures having at least one “nanoscale” dimension less than
10 300 nm, and at least one other dimension being much larger than the nanoscale dimension, such as, for example, at least about 10 or at least about 100 or at least about 200 or at least about 1000 times larger. Examples of such nanostructures are nanorods, nanowires, nanotubes, nanopyramids, nanoprisms, nanoplates, and the like. “One-dimensional” nanostructures have one dimension that is much
15 larger than the other two dimensions, such as, for example, at least about 10 or at least about 100 or at least about 200 or at least about 1000 times larger.

Such one-dimensional nanostructures may, in some cases, comprise nanowires. Nanowires are one-dimensional nanostructures in which the two short dimensions (the thickness dimensions) are less than 300 nm, preferably
20 less than 100 nm, while the third dimension (the length dimension) is greater than 1 micron, preferably greater than 10 microns, and the aspect ratio (ratio of the length dimension to the larger of the two thickness dimensions) is greater than five. Nanowires are being employed as conductors in electronic devices or as elements in optical devices, among other possible uses. Silver nanowires are
25 preferred in some such applications.

Such methods may be used to prepare nanostructures other than nanowires, such as, for example, nanocubes, nanorods, nanopyramids, nanotubes, and the like. Nanowires and other nanostructure products may be incorporated into articles, such as, for example, electronic displays, touch screens, portable
30 telephones, cellular telephones, computer displays, laptop computers, tablet computers, point-of-purchase kiosks, music players, televisions, electronic games, electronic book readers, transparent electrodes, solar cells, light emitting diodes,

other electronic devices, medical imaging devices, medical imaging media, and the like.

Preparation Methods

5 A common method of preparing nanostructures, such as, for example, nanowires, is the “polyol” process. Such a process is described in, for example, *Angew. Chem. Int. Ed.* 2009, 48, 60, Y. Xia, Y. Xiong, B. Lim, S. E. Skrabalak, which is hereby incorporated by reference in its entirety. Such processes typically reduce a metal cation, such as, for example, a silver cation, to the desired metal nanostructure product, such as, for example, a silver nanowire. Such a reduction may be carried out in a reaction mixture that may, for example, comprise one or more polyols, such as, for example, ethylene glycol (EG), propylene glycol, butanediol, glycerol, sugars, carbohydrates, and the like; one or more protecting agents, such as, for example, polyvinylpyrrolidinone (also known as polyvinylpyrrolidone or PVP), other polar polymers or copolymers, surfactants, acids, and the like; and one or more metal ions. These and other components may be used in such reaction mixtures, as is known in the art. The reduction may, for example, be carried out at one or more temperatures from about 90 °C to about 190 °C.

20

Aluminum Ions or Gallium Ions

Applicant has discovered that aluminum ions or gallium ions, such as Al^{3+} or Ga^{3+} , can be used to prepare silver nanowires. It is believed that other metal ion oxidation states, such as, for example, 1+ or 2+, may also provide useful results. It is notable that use of indium as In^{3+} , which is next heavier member of IUPAC Group 13, resulted in production of only microparticles, rather than nanowires.

25

These methods are also believed to be applicable to reducible metal cations other than silver cations, including, for example, reducible cations of other IUPAC Group 11 elements, reducible cations of other coinage metals, and the like. The methods may also be used to prepare products other than nanowires, such as, for example, nanocubes, nanorods, nanopyrramids, nanotubes, and the

30

like. Such products may be incorporated into articles, such as, for example, transparent electrodes, solar cells, light emitting diodes, other electronic devices, medical imaging devices, medical imaging media, and the like.

5

EXEMPLARY EMBODIMENTS

U.S. provisional application No. 61/488,821, filed May 23, 2011, entitled NANOWIRE PREPARATION METHODS, COMPOSITIONS, AND ARTICLES, which is hereby incorporated by reference in its entirety, disclosed the following 27 non-limiting exemplary embodiments:

10

A. A method comprising:

providing a composition comprising:

at least one first compound comprising at least one first reducible metal ion;

15

at least one second compound comprising at least one second metal or metal ion differing in atomic number from said at least one first reducible metal, said at least one second metal or metal ion comprising at least one element from IUPAC Group 13; and

at least one solvent; and

reducing the at least one first reducible metal ion to at least one first metal.

20

B. The method of embodiment A, wherein the composition further comprises at least one protecting agent.

C. The method of embodiment B, wherein the at least one protecting agent comprises at least one of: one or more surfactants, one or more acids, or one or more polar polymers.

25

D. The method of embodiment B, wherein the at least one protecting agent comprises polyvinylpyrrolidinone.

E. The method of embodiment B, further comprising inerting the at least one protecting agent.

30

F. The method of embodiment A, wherein the at least one first reducible metal ion comprises at least one coinage metal ion.

G. The method of embodiment A, wherein the at least one first reducible metal ion comprises at least one ion of an element from IUPAC Group 11.

- H. The method of embodiment A, wherein the at least one first reducible metal ion comprises at least one ion of silver.
- J. The method of embodiment A, wherein the at least one first compound comprises silver nitrate.
- 5 K. The method of embodiment A, wherein the at least one second metal or metal ion comprises gallium or an ion of gallium.
- L. The method of embodiment A, wherein the at least one second compound comprises at least one chloride.
- M. The method of embodiment A, wherein the at least one second compound
10 comprises gallium(III) chloride.
- N. The method of embodiment A, wherein the at least one solvent comprises at least one polyol.
- P. The method of embodiment A, wherein the at least one solvent comprises at least one of: ethylene glycol, propylene glycol, glycerol, one or more sugars, or
15 one or more carbohydrates.
- Q. The method of embodiment A, wherein the composition has a ratio of the total moles of the at least one second metal or metal ion to the total moles of the at least one first reducible metal ion from about 0.0001 to about 0.1.
- R. The method of embodiment A, wherein the reduction is carried out at one
20 or more temperatures from about 120 °C to about 190 °C.
- S. The method of embodiment A, further comprising inerting one or more of: the composition, the at least one compound comprising the at least one first reducible metal ion, the at least one second metal or metal ion, or the at least one solvent.
- 25 T. The at least one first metal produced according to the method of embodiment A.
- U. At least one article comprising the at least one first metal produced according to the method of embodiment A.
- V. The at least one article of embodiment U, wherein the at least one first
30 metal comprises one or more nanowires, nanocubes, nanorods, nanopyramids, or nanotubes.
- W. The at least one article of embodiment U, wherein the at least one first

metal comprises at least one object having an average diameter of between about 10 nm and about 500 nm.

X. The at least one article of embodiment U, wherein the at least one first metal comprises at least one object having an aspect ratio from about 50 to about 5 10,000.

Y. At least one metal nanowire with an average diameter of between about 10 nm and about 150 nm, and with an aspect ratio from about 50 to about 10,000.

Z. The nanowire of embodiment Y, wherein the at least one metal nanowire comprises at least one coinage metal.

10 AA. The nanowire of embodiment Y, wherein the at least one metal nanowire comprises at least one element of IUPAC Group 11.

AB. The nanowire of embodiment Y, wherein the at least one metal nanowire comprises silver.

AC. At least one article comprising the at least one metal nanowire of 15 embodiment Y.

EXAMPLES

Example 1

To a 500 mL reaction flask containing 280 mL ethylene glycol
20 (EG) was added at room temperature 1.0 g of a freshly prepared solution of 27 mM GaCl_3 in EG. This solution was stripped of at least some dissolved gases by bubbling N_2 into the solution for at 60 min using a glass pipette at room temperature with mechanical stirring while at 100 rpm. (This operation will be referred to as “degassing” the solution in the sequel.) Solutions of 0.77 M
25 polyvinylpyrrolidinone (PVP) in EG and 0.25 M AgNO_3 in EG were also degassed with nitrogen for 60 minutes, then 20 mL syringes of each were prepared. The reaction mixture was heated to 145 °C under nitrogen blanketing, then the AgNO_3 and PVP solutions were added at a constant rate over 25 minutes via a 12 gauge a TEFLON[®] fluoropolymer syringe needle. The reaction was held
30 at 145 °C for 90 minutes, and then allowed to cool to ambient temperature.

The cooled reaction mixture was diluted with an equal volume of acetone, then centrifuged at 500 G for 45 minutes. The decanted solid was then

res-dispersed in 200 mL of isopropanol, shaken for 10 minutes, then centrifuged and decanted again, followed by dilution with 15 mL isopropanol.

An optical microscope picture of the unpurified silver nanowire product is shown in Figure 1. A scanning electron microscope image of the purified silver nanowire product is shown in Figure 2. The silver nanowires had an average diameter of 52.2 ± 12.6 nm and an average length of 8.5 ± 3.1 μ m.

Example 2

To a 500 mL reaction flask was added 280 mL ethylene glycol (EG) and 1.3 g of freshly prepared 15 mM $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ in EG. This solution was degassed for at least 2 hrs using a glass pipette at room temperature with mechanical stirring while at 100 rpm. Stock solutions of 0.25 M AgNO_3 in EG and 0.77 M polyvinylpyrrolidinone (PVP) in EG were also degassed by bubbling N_2 into the solutions for 60 minutes. Two syringes were loaded with 20 mL each of the AgNO_3 and PVP solutions. The reaction mixture was heated to 155 °C under N_2 and the AgNO_3 and PVP solutions were added at a constant rate over 25 minutes via 12 gauge TEFLON[®] fluoropolymer syringe needles. The reaction was held at 155 °C for 90 minutes then allowed to cool to room temperature. From the cooled mixture, the reaction mixture was diluted by an equal volume of acetone, and centrifuged at 500G for 45 minutes. The decanted solid was re-dispersed in 200 mL isopropanol, shaken 10 minutes and centrifuged again, decanted and diluted with 15 mL isopropanol.

Figure 3 shows an optical micrograph of the silver nanowire product. The nanowires exhibited an average diameter of 50.0 ± 12.1 nm and an average length of 7.4 ± 2.8 μ m, based on measurement of at least 100 wires.

Example 3 (Comparative)

To a 500 mL reaction flask was added 280 mL ethylene glycol (EG) and 2.3 g of freshly prepared 8.5 mM $\text{InCl}_3 \cdot 4\text{H}_2\text{O}$ in EG. This solution was degassed for 2 hrs using a glass pipette at room temperature with mechanical stirring while at 100 rpm. Stock solutions of 0.25 M AgNO_3 in EG and 0.84 M polyvinylpyrrolidinone (PVP) in EG were also degassed by bubbling N_2 into the

solutions for at least 60 minutes. Two syringes were loaded with 20 mL each of the AgNO_3 and PVP solutions. The reaction mixture was heated to 155 °C under N_2 and the AgNO_3 and PVP solutions were added at a constant rate over 25 minutes via 12 gauge TEFLON[®] fluoropolymer syringe needles. The reaction
5 was held at 155 °C for 90 minutes then allowed to cool to room temperature.

No nanowires were found in the reaction product, only microparticles.

CLAIMS:

1. A method comprising:
providing at least one second metal or metal ion comprising at least one of an aluminum ion or a gallium ion; and
5 reducing at least one first reducible metal ion to at least one first metal nanowire,
wherein the at least one second metal or metal ion and the at least one first reducible metal ion differ in atomic number.
- 10 2. The method according to claim 1, wherein the providing the at least one second metal or metal ion comprises providing a composition comprising:
at least one first compound comprising the at least one first reducible metal ion;
at least one second compound comprising the at least one second
15 metal or metal ion; and
at least one solvent.
3. The method of claim 2, wherein the at least one second compound comprises at least one chloride ion.
20
4. The method of claim 2, wherein the at least one second compound comprises gallium(III) chloride.
5. The method of claim 2, wherein the at least one second compound
25 comprises aluminum(III) chloride hexahydrate.
6. The method of claim 1, wherein the reducing the at least one first reducible metal ion occurs in the presence of the at least one second metal or metal ion.
30

7. The method of claim 1, wherein the at least one first reducible metal ion comprises at least of a coinage metal ion or an ion of an element from IUPAC Group 11 or at least one ion of silver.
- 5 8. The method of claim 1, wherein the at least one second metal or metal ion comprises gallium or an ion of gallium.
9. The method of claim 1, wherein the at least one second metal or metal ion comprises aluminum or an ion of aluminum.
- 10 10. The at least one first metal nanowire produced according to the method of embodiment 1.

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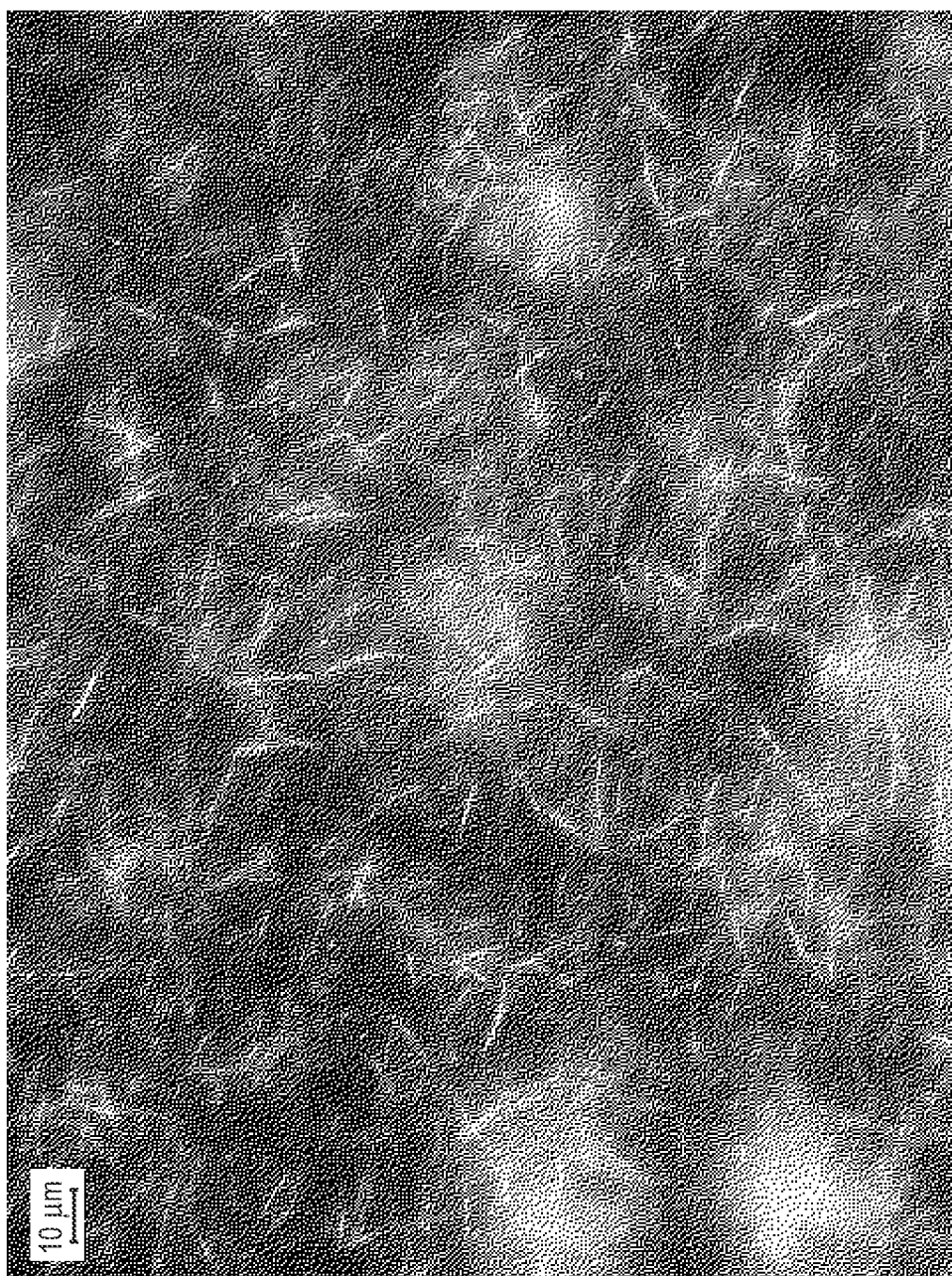
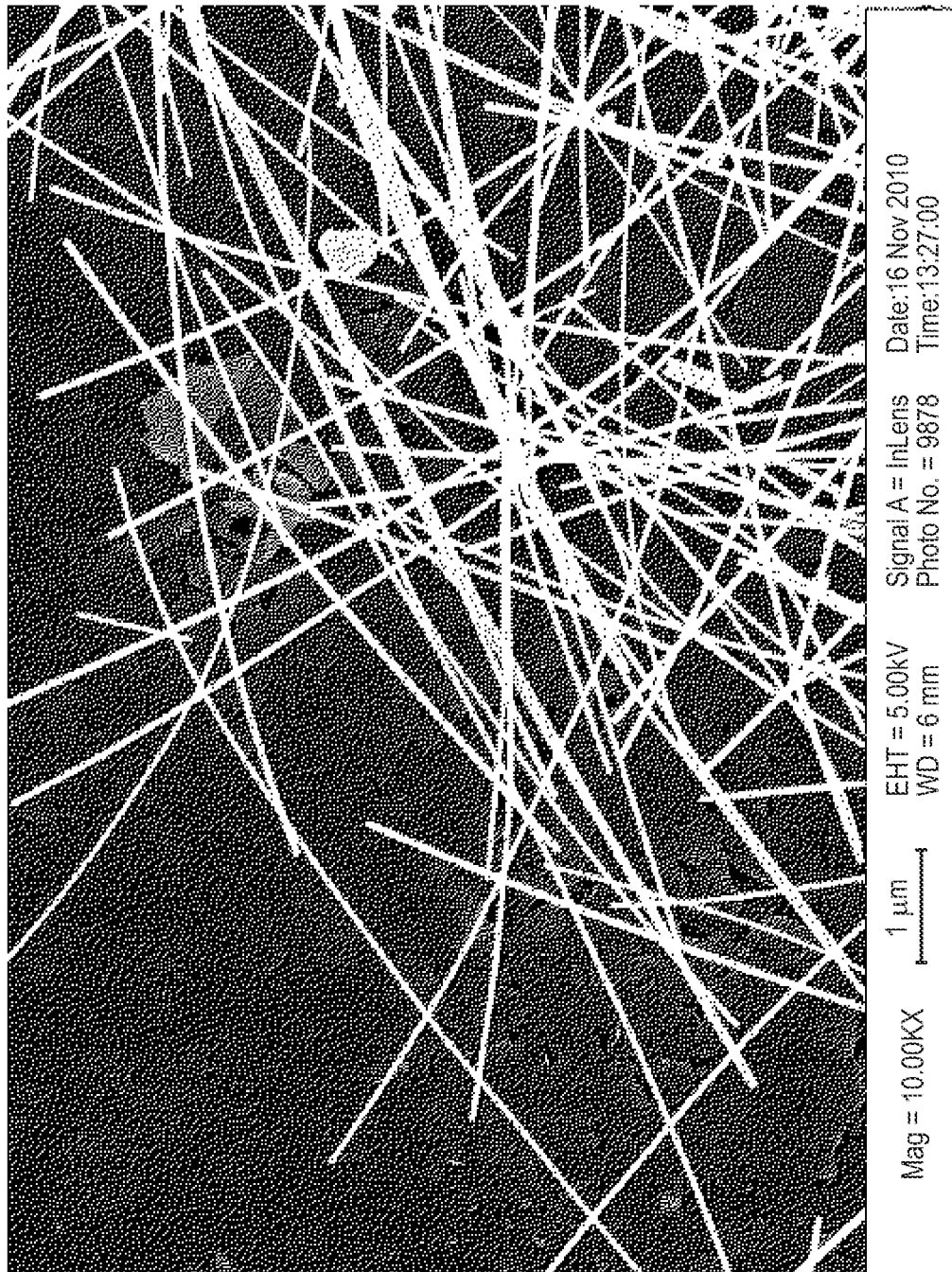


FIG. 1

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**FIG. 2**

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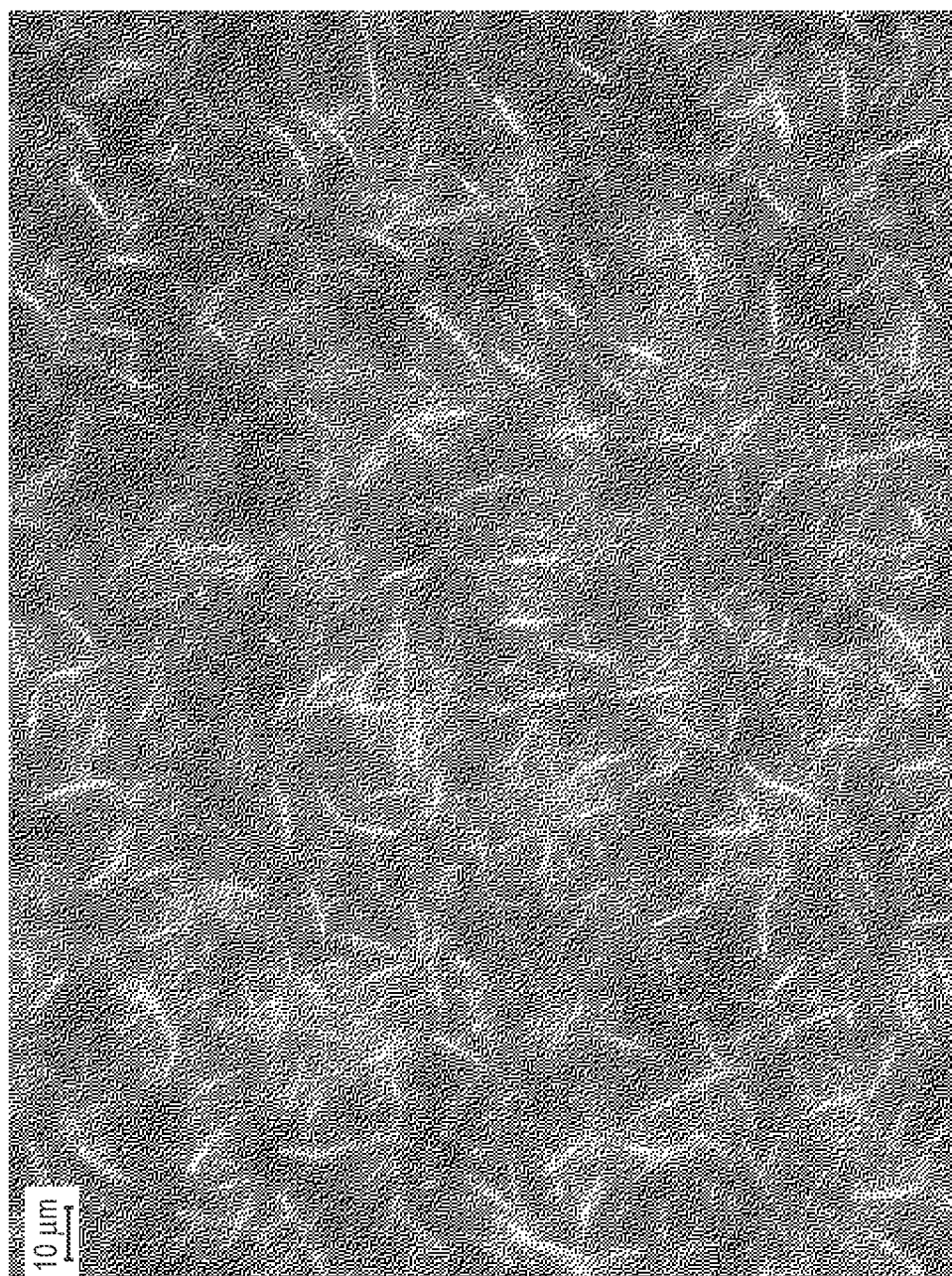


FIG. 3

INTERNATIONAL SEARCH REPORT

International application No
PCT/US2012/034141

A. CLASSIFICATION OF SUBJECT MATTER

INV. B22F9/24 C30B29/60 C30B7/14
ADD.

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

B22F C30B

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

EPO-Internal, COMPENDEX, INSPEC, WPI Data

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	KYLEE E. KORTE ET AL: "Rapid synthesis of silver nanowires through a CuCl- or CuCl2-mediated polyol process", JOURNAL OF MATERIALS CHEMISTRY, vol. 18, no. 4, 1 January 2008 (2008-01-01), page 437, XP55010797, ISSN: 0959-9428, DOI: 10.1039/b714072j cited in the application	10
A	Experimental ----- -/--	1-9

☒ Further documents are listed in the continuation of Box C.

☐ See patent family annex.

* Special categories of cited documents :

"A" document defining the general state of the art which is not considered to be of particular relevance

"E" earlier application or patent but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

"&" document member of the same patent family

Date of the actual completion of the international search

3 July 2012

Date of mailing of the international search report

10/07/2012

Name and mailing address of the ISA/

European Patent Office, P.B. 5818 Patentlaan 2
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Authorized officer

Morra, Valentina

INTERNATIONAL SEARCH REPORT

International application No
PCT/US2012/034141

C(Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	HONG-YAN SHI 1 ET AL: "Ordering of Disordered Nanowires: Spontaneous Formation of Highly Aligned, Ultralong Ag Nanowire Films at Oil-Water-Air Interface", ADVANCED FUNCTIONAL MATERIALS, WILEY - V C H VERLAG GMBH & CO. KGAA, DE, vol. 20, no. 6, 24 March 2010 (2010-03-24), pages 958-964, XP001554277, ISSN: 1616-301X, DOI: 10.1002/ADFM.200901668	10
A	Experimental -----	1-9

INTERNATIONAL SEARCH REPORT

International application No.
PCT/US2012/034141

Box No. II Observations where certain claims were found unsearchable (Continuation of item 2 of first sheet)

This international search report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☐ Claims Nos.:
because they relate to subject matter not required to be searched by this Authority, namely:
2. ☒ Claims Nos.: 1-10(partially)
because they relate to parts of the international application that do not comply with the prescribed requirements to such an extent that no meaningful international search can be carried out, specifically:
see FURTHER INFORMATION sheet PCT/ISA/210
3. ☐ Claims Nos.:
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box No. III Observations where unity of invention is lacking (Continuation of item 3 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

1. ☐ As all required additional search fees were timely paid by the applicant, this international search report covers all searchable claims.
2. ☐ As all searchable claims could be searched without effort justifying an additional fees, this Authority did not invite payment of additional fees.
3. ☐ As only some of the required additional search fees were timely paid by the applicant, this international search report covers only those claims for which fees were paid, specifically claims Nos.:
4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this international search report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest and, where applicable, the payment of a protest fee.
- ☐ The additional search fees were accompanied by the applicant's protest but the applicable protest fee was not paid within the time limit specified in the invitation.
- ☐ No protest accompanied the payment of additional search fees.

FURTHER INFORMATION CONTINUED FROM PCT/ISA/ 210

Continuation of Box II.2

Claims Nos.: 1-10(partially)

Present claims 1-10 relate to an extremely large number of possible methods and products. Support and disclosure in the sense of Article 6 and 5 PCT is to be found however for only a very small proportion of the methods and products claimed, i.e. for methods of producing silver nanowires, see Examples. The non-compliance with the substantive provisions is to such an extent, that the search was performed taking into consideration the non-compliance in determining the extent of the search of claims 1-10 (PCT Guidelines 9.19 and 9.23). The search of claims 1-10 was restricted to those claimed methods and products which appear to be supported and a generalisation of their composition, i.e. to methods of producing Ag, Au or Cu nanowires.

The applicant's attention is drawn to the fact that claims relating to inventions in respect of which no international search report has been established need not be the subject of an international preliminary examination (Rule 66.1(e) PCT). The applicant is advised that the EPO policy when acting as an International Preliminary Examining Authority is normally not to carry out a preliminary examination on matter which has not been searched. This is the case irrespective of whether or not the claims are amended following receipt of the search report or during any Chapter II procedure. If the application proceeds into the regional phase before the EPO, the applicant is reminded that a search may be carried out during examination before the EPO (see EPO Guideline C-VI, 8.2), should the problems which led to the Article 17(2) declaration be overcome.