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(54) Title: METHODS FOR DETECTING VITAMIN D METABOLITES BY MASS SPECTROMETRY

(57) Abstract: Provided are methods of detecting the presence or amount of a vitamin D metabolite in a sample using mass spectrometry. The methods generally comprise ionizing a vitamin D metabolite in a sample and detecting the amount of the ion to determine the presence or amount of the vitamin D metabolite in the sample. Also provided are methods to detect the presence or amount of two or more vitamin D metabolites in a single assay.

METHODS FOR DETECTING VITAMIN D METABOLITES BY MASS SPECTROMETRY

FIELD OF THE INVENTION

[0001] The invention relates to the detection of vitamin D metabolites. In a particular aspect, the invention relates to methods for detecting vitamin D metabolites by mass spectrometry.

BACKGROUND OF THE INVENTION

[0002] Vitamin D is an essential nutrient with important physiological roles in the positive regulation of calcium (Ca^{2+}) homeostasis. Vitamin D can be made *de novo* in the skin by exposure to sunlight or it can be absorbed from the diet. There are two forms of vitamin D; vitamin D₂ (ergocalciferol) and vitamin D₃ (cholecalciferol). Vitamin D₃ is the form synthesized *de novo* by animals. It is also a common supplement added to milk products and certain food products produced in the United States. Both dietary and intrinsically synthesized vitamin D₃ must undergo metabolic activation to generate bioactive metabolites. In humans, the initial step of vitamin D₃ activation occurs primarily in the liver and involves hydroxylation to form the intermediate metabolite 25-hydroxyvitamin D₃ (25-hydroxycholecalciferol; calcifediol; 25OHD₃). Calcifediol is the major form of vitamin D₃ in the circulation. Circulating 25OHD₃ is then converted by the kidney to 1,25-dihydroxyvitamin D₃ (calcitriol; 1,25(OH)₂D₃), which is generally believed to be the metabolite of vitamin D₃ with the highest biological activity.

[0003] Vitamin D₂ is derived from fungal and plant sources. Many over-the-counter dietary supplements contain ergocalciferol (vitamin D₂) rather than cholecalciferol (vitamin D₃).

Drisdol, the only high-potency prescription form of vitamin D available in the United States, is formulated with ergocalciferol. Vitamin D₂ undergoes a similar pathway of metabolic activation in humans as vitamin D₃, forming the metabolites 25-hydroxyvitamin D₂ (25OHD₂) and 1,25-dihydroxyvitamin D₃ (1,25(OH)₂D₂). Vitamin D₂ and vitamin D₃ have long been assumed to be biologically equivalent in humans, however recent reports suggest that there may be differences in the bioactivity and bioavailability of these two forms of vitamin D (Armas *et. al.*, (2004) J. Clin. Endocrinol. Metab. 89:5387-5391).

[0004] Measurement of vitamin D, the inactive vitamin D precursor, is rare in clinical settings and has little diagnostic value. Rather, serum levels of 25-hydroxyvitamin D₃ and 25-hydroxyvitamin D₂ (total 25-hydroxyvitamin D; "25OHD") are a useful index of vitamin D nutritional status and the efficacy of certain vitamin D analogs. Therefore, the measurement of 25OHD is commonly used in the diagnosis and management of disorders of calcium metabolism. In this respect, low levels of 25OHD are indicative of vitamin D deficiency associated with diseases such as hypocalcemia, hypophosphatemia, secondary hyperparathyroidism, elevated alkaline phosphatase, osteomalacia in adults and rickets in children. In patients suspected of vitamin D intoxication, elevated levels of 25OHD distinguishes this disorder from other disorders that cause hypercalcemia.

[0005] Measurement of 1,25(OH)₂D is also used in clinical settings, however, this metabolite has a more limited diagnostic usefulness than 25OHD. Factors that contribute to limitations of the diagnostic values of 1,25(OH)₂D as an index of vitamin D status include the precision of the endogenous regulation of renal production of the metabolite and its short half-life in circulation. However, certain disease states such as kidney failure can be diagnosed by reduced levels of circulating 1,25(OH)₂D and elevated levels of 1,25(OH)₂D may be indicative of excess parathyroid hormone or may be indicative of certain diseases such as sarcoidosis or certain types of lymphoma.

[0006] Detection of vitamin D metabolites has been accomplished by radioimmunoassay with antibodies co-specific for 25-hydroxyvitamin D₃ and 25-hydroxyvitamin D₂. Because the current immunologically-based assays do not separately resolve 25-hydroxyvitamin D₃ and 25-hydroxyvitamin D₂, the source of a deficiency in vitamin D nutrition cannot be determined without resorting to other tests. More recently, reports have been published that disclose methods for detecting specific vitamin D metabolites using mass spectrometry. For

example Yeung B, *et al.*, J Chromatogr. 1993, 645(1):115-23; Higashi T, *et al.*, Steroids. 2000, 65(5):281-94; Higashi T, *et al.*, Biol Pharm Bull. 2001, 24(7):738-43; and Higashi T, *et al.*, J Pharm Biomed Anal. 2002, 29(5):947-55 disclose methods for detecting various vitamin D metabolites using liquid chromatography and mass spectrometry. These methods require that the metabolites be derivatized prior to detection by mass-spectrometry. Methods to detect underivatized 1,25(OH)₂D₃ by liquid chromatography / mass-spectrometry are disclosed in Kissmeyer and Sonne, J Chromatogr A. 2001, 935(1-2):93-103.

SUMMARY OF THE INVENTION

[0007] The present invention provides methods for detecting the presence or amount of a vitamin D metabolite in a sample by mass spectrometry, including tandem mass spectrometry. Preferably, the methods of the invention do not include derivatizing the vitamin D metabolites prior to the mass spectrometry analysis.

[0008] In one aspect, the invention provides a method for determining the presence or amount of a vitamin D metabolite in a sample. The method may include: (a) ionizing the vitamin D metabolite, if present in the sample; and (b) detecting the presence or amount of the ion by mass spectrometry and relating presence or amount of the detected ion to the presence or amount of the vitamin D metabolite in the sample. In some preferred embodiments, the ionization step (a) may include (i) ionizing the vitamin D metabolite, if present in the sample, to produce an ion of the vitamin D metabolite; (ii) isolating the ion of step (i) by mass spectrometry to provide a precursor ion; and (iii) effecting a collision between the precursor ion and an inert collision gas to produce at least one fragment ion detectable in a mass spectrometer. Preferably, at least one of the fragment ions is specific for the vitamin D metabolite of interest. In certain embodiments of the invention, the fragment ions to be detected include at least one fragment ion other than that which results solely by a dehydration or deamination of the precursor ion. In some particularly preferred embodiments, the precursor ion is a protonated and dehydrated ion of the vitamin D metabolite. In certain embodiments the vitamin D metabolite is one or more vitamin D metabolites selected from the group consisting of 25-hydroxyvitamin D₃; 25-hydroxyvitamin D₂; 1,25-dihydroxyvitamin D₂; and 1,25-dihydroxyvitamin D₃.

[0009] In a related aspect, the invention provides a method for determining the presence or amount of a vitamin D metabolite in a sample by tandem mass spectrometry. The method may involve (a) generating a protonated and dehydrated precursor ion of the vitamin D metabolite if present in the sample; (b) generating one or more fragment ions of the precursor ion; and (c) detecting the presence or amount of one or more of the ions generated in step (a) or (b) or both and relating the detected ions to the presence or amount of the vitamin D metabolite in the sample. In certain embodiments, the method is used to detect the presence or amount of two or more vitamin D metabolites in a single assay. Preferably, the method does not involve derivatizing the samples or the vitamin D metabolites prior to analysis by mass spectrometry. In certain embodiments the vitamin D metabolite is one or more vitamin D metabolites selected from the group consisting of 25-hydroxyvitamin D₃; 25-hydroxyvitamin D₂; 1,25-dihydroxyvitamin D₂; and 1,25-dihydroxyvitamin D₃.

[0010] In another aspect the invention provides a method for determining the presence or amount of two or more vitamin D metabolites in a sample in a single assay. The method includes ionizing the vitamin D metabolites, if present in the sample, to generate ions specific for each of the vitamin D metabolites of interest, detecting the presence or amount of the ions by mass spectrometry, and relating the presence or amount of the ions to the presence or amount of the vitamin D metabolites in the sample. In certain embodiments the mass spectrometry analysis of the method is tandem mass spectrometry.

[0011] As used herein, the term "vitamin D metabolite" refers to any chemical species that may be found in the circulation of an animal which is formed by a biosynthetic or metabolic pathway for vitamin D or a synthetic vitamin D analog. Vitamin D metabolites include forms of vitamin D that are generated by a biological organism, such as an animal, or that are generated by biotransformation of a naturally occurring form of vitamin D or a synthetic vitamin D analog. In certain preferred embodiments, a vitamin D metabolite is formed by the biotransformation of vitamin D₂ or vitamin D₃. In particularly preferred embodiments, the vitamin D metabolite is one or more compounds selected from the group consisting of 25-hydroxyvitamin D₃, 25-hydroxyvitamin D₂, 1,25-dihydroxyvitamin D₃ and 1,25-dihydroxyvitamin D₂.

[0012] As used herein, the term "purification" refers to a procedure that enriches the amount of one or more analytes of interest relative to one or more other components of the sample.

Purification, as used herein does not require the isolation of an analyte from all others. In preferred embodiments, a purification step or procedure can be used to remove one or more interfering substances, *e.g.*, one or more substances that would interfere with the operation of the instruments used in the methods or substances that may interfere with the detection of an analyte ion by mass spectrometry.

[0013] As used herein, “biological sample” refers to any sample from a biological source. As used herein, “body fluid” means any fluid that can be isolated from the body of an individual. For example, “body fluid” may include blood, plasma, serum, bile, saliva, urine, tears, perspiration, and the like.

[0014] As used herein, “derivatizing” means reacting two molecules to form a new molecule. Derivatizing agents may include isothiocyanate groups, dinitro-fluorophenyl groups, nitrophenoxycarbonyl groups, and/or phthalaldehyde groups.

[0015] As used herein, “chromatography” refers to a process in which a chemical mixture carried by a liquid or gas is separated into components as a result of differential distribution of the chemical entities as they flow around or over a stationary liquid or solid phase.

[0016] As used herein, “liquid chromatography” (LC) means a process of selective retardation of one or more components of a fluid solution as the fluid uniformly percolates through a column of a finely divided substance, or through capillary passageways. The retardation results from the distribution of the components of the mixture between one or more stationary phases and the bulk fluid, (*i.e.*, mobile phase), as this fluid moves relative to the stationary phase(s). “Liquid chromatography” includes reverse phase liquid chromatography (RPLC), high performance liquid chromatography (HPLC) and high turbulence liquid chromatography (HTLC).

[0017] As used herein, the term “HPLC” or “high performance liquid chromatography” refers to liquid chromatography in which the degree of separation is increased by forcing the mobile phase under pressure through a stationary phase, typically a densely packed column.

[0018] As used herein, the term “gas chromatography” refers to chromatography in which the sample mixture is vaporized and injected into a stream of carrier gas (as nitrogen or helium) moving through a column containing a stationary phase composed of a liquid or a

particulate solid and is separated into its component compounds according to the affinity of the compounds for the stationary phase

[0019] As used herein, “mass spectrometry” (MS) refers to an analytical technique to identify compounds by their mass. MS technology generally includes (1) ionizing the compounds to form charged compounds; and (2) detecting the molecular weight of the charged compound and calculating a mass-to-charge ratio (m/z). The compound may be ionized and detected by any suitable means. A “mass spectrometer” generally includes an ionizer and an ion detector. *See, e.g.*, U.S. Patent Nos. 6,204,500, entitled “Mass Spectrometry From Surfaces;” 6,107,623, entitled “Methods and Apparatus for Tandem Mass Spectrometry;” 6,268,144, entitled “DNA Diagnostics Based On Mass Spectrometry;” 6,124,137, entitled “Surface-Enhanced Photolabile Attachment And Release For Desorption And Detection Of Analytes;” Wright *et al.*, *Prostate Cancer and Prostatic Diseases* 2:264-76 (1999); and Merchant and Weinberger, *Electrophoresis* 21:1164-67 (2000).

[0020] The term “electron ionization” as used herein refers to methods in which an analyte of interest in a gaseous or vapor phase interacts with a flow of electrons. Impact of the electrons with the analyte produces analyte ions, which may then be subjected to a mass spectrometry technique.

[0021] The term “chemical ionization” as used herein refers to methods in which a reagent gas (*e.g.* ammonia) is subjected to electron impact, and analyte ions are formed by the interaction of reagent gas ions and analyte molecules.

[0022] The term “fast atom bombardment” as used herein refers to methods in which a beam of high energy atoms (often Xe or Ar) impacts a non-volatile sample, desorbing and ionizing molecules contained in the sample. Test samples are dissolved in a viscous liquid matrix such as glycerol, thioglycerol, m-nitrobenzyl alcohol, 18-crown-6 crown ether, 2-nitrophenyloctyl ether, sulfolane, diethanolamine, and triethanolamine.

[0023] The term “field desorption” as used herein refers to methods in which a non-volatile test sample is placed on an ionization surface, and an intense electric field is used to generate analyte ions.

[0024] The term “ionization” as used herein refers to the process of generating an analyte ion having a net electrical charge equal to one or more electron units. Negative ions are those

having a net negative charge of one or more electron units, while positive ions are those having a net positive charge of one or more electron units.

[0025] The term “operating in negative ion mode” refers to those mass spectrometry methods where negative ions are detected. Similarly, “operating in positive ion mode” refers to those mass spectrometry methods where positive ions are detected.

[0026] The term “desorption” as used herein refers to the removal of an analyte from a surface and/or the entry of an analyte into a gaseous phase.

[0027] The term “about” as used herein in reference to quantitative measurements, refers to the indicated value plus or minus 10%.

BRIEF DESCRIPTION OF THE DRAWINGS

[0028] Figure 1 shows the linearity of the quantification of 25OHD₂ in serially diluted stock samples using an LC-MS/MS assay. Details are described in Example 4.

[0029] Figure 2 shows the linearity of the quantification of 25OHD₃ in serially diluted stock samples using an LC-MS/MS assay. Details are described in Example 4.

[0030] Figure 3 shows the linearity of the quantification by LC-MS/MS of serially diluted samples containing 25OHD₂ and 25OHD₃ to final concentrations of 512 ng/mL. Details are described in Example 4.

[0031] Figure 4 shows the correlation between detection of total 25-hydroxyvitamin D using an LC-MS/MS assay and a commercially available radioimmunoassay kit. Details are described in Example 6.

DETAILED DESCRIPTION OF THE INVENTION

[0032] Disclosed are methods for detecting the presence or amount of one or more vitamin D metabolites in a sample. In certain aspects the method involves ionizing the vitamin D metabolite(s), detecting the ion(s) by mass spectrometry, and relating the presence or amount of the ion(s) to the presence or amount of the vitamin D metabolite(s) in the sample. The method may include (a) purifying a vitamin D metabolite, if present in the sample, (b)

ionizing the purified vitamin D metabolite and (c) detecting the presence or amount of the ion, wherein the presence or amount of the ion is related to the presence or amount of the vitamin D metabolite in the sample. In preferred embodiments, the ionizing step (b) may include (i) ionizing a vitamin D metabolite, if present in the sample, to produce an ion; (ii) isolating the vitamin D metabolite ion by mass spectrometry to provide a precursor ion; and (iii) effecting a collision between the isolated precursor ion and an inert collision gas to produce at least one fragment ion detectable in a mass spectrometer. In certain preferred embodiments the precursor ion is a protonated and dehydrated ion of the vitamin D metabolite.

[0033] In a related aspect, the invention provides a method for determining the presence or amount of a vitamin D metabolite in a test sample by tandem mass spectrometry. The method may involve (a) generating a protonated and dehydrated precursor ion of the vitamin D metabolite; (b) generating one or more fragment ions of the precursor ion; and (c) detecting the presence or amount of one or more of the ions generated in step (a) or (b) or both and relating the detected ions to the presence or amount of said vitamin D metabolite in the sample.

[0034] In certain preferred embodiments of the invention, at least one fragment ion is detected, wherein the presence or amount of the precursor and/or at least one fragment ion is related to the presence or amount of the vitamin D metabolite in the sample. Preferably at least one fragment ion is specific for the vitamin D metabolite of interest. In some embodiments, the methods of the invention can be used to detect and quantify two or more vitamin D metabolites in a single assay. In certain embodiments, the vitamin D metabolite is one or more vitamin D metabolites selected from the group consisting of 25-hydroxyvitamin D₃; 25-hydroxyvitamin D₂; 1,25-dihydroxyvitamin D₂; and 1,25-dihydroxyvitamin D₃.

[0035] Suitable samples include any sample that might contain the analyte of interest and/or one or more metabolites or precursors thereof. For example, samples obtained during the manufacture of an analyte can be analyzed to determine the composition and yield of the manufacturing process. In certain embodiments, a sample is a biological sample; that is, a sample obtained from any biological source, such as an animal, a cell culture, an organ culture, *etc.* Particularly preferred are samples obtained from a human, such as a blood, plasma, serum, hair, muscle, urine, saliva, tear, cerebrospinal fluid, or other tissue sample.

Such samples may be obtained, for example, from a patient seeking diagnosis, prognosis, or treatment of a disease or condition. The vitamin D metabolites may be derivatized prior to mass spectrometry, however, in certain preferred embodiments, sample preparation excludes the use of derivitization.

[0036] Samples may be processed or purified to obtain preparations that are suitable for analysis by mass spectrometry. Such purification will usually include chromatography, such as liquid chromatography, and may also often involve an additional purification procedure that is performed prior to chromatography. Various procedures may be used for this purpose depending on the type of sample or the type of chromatography. Examples include filtration, extraction, precipitation, centrifugation, dilution, combinations thereof and the like. Protein precipitation is one preferred method of preparing a liquid biological sample, such as serum or plasma, for chromatography. Such protein purification methods are well known in the art, for example, Polson *et al.*, *Journal of Chromatography B* **785**:263-275 (2003), describes protein precipitation methods suitable for use in the methods of the invention. Protein precipitation may be used to remove most of the protein from the sample leaving vitamin D metabolites soluble in the supernatant. The samples can be centrifuged to separate the liquid supernatant from the precipitated proteins. The resultant supernatant can then be applied to liquid chromatography and subsequent mass spectrometry analysis. In one embodiment of the invention, the protein precipitation involves adding one volume of the liquid sample (e.g. plasma) to about four volumes of methanol. In certain embodiments, the use of protein precipitation obviates the need for high turbulence liquid chromatography (“HTLC”) or on-line extraction prior to HPLC and mass spectrometry. Accordingly in such embodiments, the method involves (1) performing a protein precipitation of the sample of interest; and (2) loading the supernatant directly onto the HPLC-mass spectrometer without using on-line extraction or high turbulence liquid chromatography (“HTLC”).

[0037] The purification step may include chromatography, preferably liquid chromatography, more preferably high performance liquid chromatography (HPLC). In some preferred embodiments the chromatography is not gas chromatography. Preferably, the methods of the invention are performed without subjecting the samples, or the vitamin D metabolites of interest, to gas chromatography prior to mass spectrometric analysis.

[0038] Various methods have been described involving the use of HPLC for sample clean-up prior to mass spectrometry analysis. See, *e.g.*, Taylor *et al.*, *Therapeutic Drug Monitoring* **22**:608-12 (2000) (manual precipitation of blood samples, followed by manual C18 solid phase extraction, injection into an HPLC for chromatography on a C18 analytical column, and MS/MS analysis); and Salm *et al.*, *Clin. Therapeutics* **22** Supl. B:B71-B85 (2000) (manual precipitation of blood samples, followed by manual C18 solid phase extraction, injection into an HPLC for chromatography on a C18 analytical column, and MS/MS analysis). One of skill in the art can select HPLC instruments and columns that are suitable for use in the invention. The chromatographic column typically includes a medium (*i.e.*, a packing material) to facilitate separation of chemical moieties (*i.e.*, fractionation). The medium may include minute particles. The particles include a bonded surface that interacts with the various chemical moieties to facilitate separation of the chemical moieties such as vitamin D metabolites. One suitable bonded surface is a hydrophobic bonded surface such as an alkyl bonded surface. Alkyl bonded surfaces may include C-4, C-8, or C-18 bonded alkyl groups, preferably C-18 bonded groups. The chromatographic column includes an inlet port for receiving a sample and an outlet port for discharging an effluent that includes the fractionated sample. In the method, the sample (or pre-purified sample) is applied to the column at the inlet port, eluted with a solvent or solvent mixture, and discharged at the outlet port. Different solvent modes may be selected for eluting the analytes of interest. For example, liquid chromatography may be performed using a gradient mode, an isocratic mode, or a polytypic (*i.e.* mixed) mode. In preferred embodiments, HPLC is performed on a multiplexed analytical HPLC system with a C18 solid phase using isocratic separation with 100% methanol as the mobile phase.

[0039] Recently, high turbulence liquid chromatography (“HTLC”), also called high throughput liquid chromatography, has been applied for sample preparation prior to analysis by mass spectrometry. See, *e.g.*, Zimmer *et al.*, *J. Chromatogr. A* **854**:23-35 (1999); see also, U.S. Patents Nos. 5,968,367; 5,919,368; 5,795,469; and 5,772,874. Traditional HPLC analysis relies on column packings in which laminar flow of the sample through the column is the basis for separation of the analyte of interest from the sample. The skilled artisan will understand that separation in such columns is a diffusional process. In contrast, it is believed that turbulent flow, such as that provided by HTLC columns and methods, may enhance the rate of mass transfer, improving the separation characteristics provided. In some

embodiments, high turbulence liquid chromatography (HTLC), alone or in combination with one or more purification methods, may be used to purify the vitamin D metabolite of interest prior to mass spectrometry. In such embodiments samples may be extracted using an HTLC extraction cartridge which captures the analyte, then eluted and chromatographed on a second HTLC column or onto an analytical HPLC column prior to ionization. Because the steps involved in these chromatography procedures can be linked in an automated fashion, the requirement for operator involvement during the purification of the analyte can be minimized. In certain embodiments of the method, samples are subjected to protein precipitation as described above prior to loading on the HTLC column; in alternative embodiments, the samples may be loaded directly onto the HTLC without being subjected to protein precipitation.

[0040] Recently, research has shown that epimerization of the hydroxyl group of the A-ring of vitamin D₃ metabolites is an important aspect of vitamin D₃ metabolism and bioactivation, and that depending on the cell types involved, 3-C epimers of vitamin D₃ metabolites (e.g., 3-epi-25(OH)D₃; 3-epi-24,25(OH)₂D₃; and 3-epi-1,25(OH)₂D₃) are often major metabolic products. See Kamao *et al.*, *J. Biol. Chem.*, **279**:15897-15907 (2004). Kamao *et al.*, further provides methods of separating various vitamin D metabolites, including 3-C epimers, using Chiral HPLC. Accordingly, the invention also provides methods of detecting the presence, absence and/or amount of a specific epimer of one or more vitamin D metabolites, preferably vitamin D₃ metabolites, in a sample by (1) separating one or more specific vitamin D metabolites by chiral chromatography, preferably chiral HPLC; and (2) detecting the presence and/or amount of one or more vitamin D metabolites using mass spectrometry methods as described herein. The chiral chromatography procedures described in Kamao *et al.*, are suitable for the methods of the invention, however, one of ordinary skill in the art understands that there are numerous other chiral chromatography methods that would also be suitable. In preferred embodiments the method includes, separating 25(OH)D₃ from 3-epi-25(OH)D₃, if present in a sample, using chiral chromatography; and detecting the presence and/or amount of the 25(OH)D₃ and the 3-epi-25(OH)D₃ in the sample using mass spectrometry. In related embodiments, the method includes separating 1 α ,25(OH)₂D₃ from 3-epi-1 α ,25(OH)₂D₃, if present in a sample, using chiral chromatography; and detecting the presence and/or amount of the 1 α ,25(OH)₂D₃ and the 3-epi-1 α ,25(OH)₂D₃ in the sample

using mass spectrometry. In certain embodiments of the invention, chiral chromatography is used in conjunction with the HTLC methods described above.

[0041] Mass spectrometry is performed using a mass spectrometer which includes an ion source for ionizing the fractionated sample and creating charged molecules for further analysis. For example ionization of the sample may be performed by electrospray ionization (ESI), atmospheric pressure chemical ionization (APCI), photoionization, electron ionization, fast atom bombardment (FAB)/liquid secondary ionization (LSIMS), matrix assisted laser desorption ionization (MALDI), field ionization, field desorption, thermospray/plasmaspray ionization, and particle beam ionization. The skilled artisan will understand that the choice of ionization method can be determined based on the analyte to be measured, type of sample, the type of detector, the choice of positive versus negative mode, *etc.*

[0042] After the sample has been ionized, the positively charged or negatively charged ions thereby created may be analyzed to determine a mass-to-charge ratio (*i.e.*, m/z). Suitable analyzers for determining mass-to-charge ratios include quadropole analyzers, ion traps analyzers, and time-of-flight analyzers. The ions may be detected using several detection modes. For example, selected ions may be detected (*i.e.*, using a selective ion monitoring mode (SIM)), or alternatively, ions may be detected using a scanning mode, e.g., multiple reaction monitoring (MRM) or selected reaction monitoring (SRM). Preferably, the mass-to-charge ratio is determined using a quadropole analyzer. For example, in a “quadropole” or “quadropole ion trap” instrument, ions in an oscillating radio frequency field experience a force proportional to the DC potential applied between electrodes, the amplitude of the RF signal, and m/z . The voltage and amplitude can be selected so that only ions having a particular m/z travel the length of the quadropole, while all other ions are deflected. Thus, quadropole instruments can act as both a “mass filter” and as a “mass detector” for the ions injected into the instrument.

[0043] One can often enhance the resolution of the MS technique by employing “tandem mass spectrometry,” or “MS/MS.” In this technique, a precursor ion (also called a parent ion) generated from a molecule of interest can be filtered in an MS instrument, and the precursor ion is subsequently fragmented to yield one or more fragment ions (also called daughter ions or product ions) that are then analyzed in a second MS procedure. By careful selection of

precursor ions, only ions produced by certain analytes are passed to the fragmentation chamber, where collision with atoms of an inert gas to produce the daughter ions. Because both the precursor and fragment ions are produced in a reproducible fashion under a given set of ionization/fragmentation conditions, the MS/MS technique can provide an extremely powerful analytical tool. For example, the combination of filtration/fragmentation can be used to eliminate interfering substances, and can be particularly useful in complex samples, such as biological samples.

[0044] Additionally, recent advances in technology, such as matrix-assisted laser desorption ionization coupled with time-of-flight analyzers (“MALDI-TOF”) permit the analysis of analytes at femtomole levels in very short ion pulses. Mass spectrometers that combine time-of-flight analyzers with tandem MS are also well known to the artisan. Additionally, multiple mass spectrometry steps can be combined in methods known as “MS/MSⁿ.” Various other combinations may be employed, such as MS/MS/TOF, MALDI/MS/MS/TOF, or SELDI/MS/MS/TOF mass spectrometry.

[0045] The mass spectrometer typically provides the user with an ion scan; that is, the relative abundance of each ion with a particular m/z over a given range (*e.g.*, 100 to 1000 amu). The results of an analyte assay, that is, a mass spectrum, can be related to the amount of the analyte in the original sample by numerous methods known in the art. For example, given that sampling and analysis parameters are carefully controlled, the relative abundance of a given ion can be compared to a table that converts that relative abundance to an absolute amount of the original molecule. Alternatively, molecular standards can be run with the samples, and a standard curve constructed based on ions generated from those standards. Using such a standard curve, the relative abundance of a given ion can be converted into an absolute amount of the original molecule. In certain preferred embodiments, an internal standard is used to generate a standard curve for calculating the quantity of the vitamin D metabolite. Methods of generating and using such standard curves are well known in the art and one of ordinary skill is capable of selecting an appropriate internal standard. For example, an isotope of a vitamin D metabolite may be used as an internal standard, in preferred embodiments the vitamin D metabolite is a deuterated vitamin D metabolite, for example ⁶D-25OHD₃. Numerous other methods for relating the presence or amount of an ion

to the presence or amount of the original molecule will be well known to those of ordinary skill in the art.

[0046] One or more steps of the methods of the invention can be performed using automated machines. In certain embodiments, one or more purification steps are performed on line, and more preferably all of the purification and mass spectrometry steps may be performed in an on-line fashion.

[0047] In particularly preferred embodiments vitamin D metabolites are detected and/or quantified using LC-MS/MS as follows. The samples are subjected to liquid chromatography, preferably HPLC, the flow of liquid solvent from the chromatographic column enters the heated nebulizer interface of a LC-MS/MS analyzer and the solvent/analyte mixture is converted to vapor in the heated tubing of the interface. The analytes (i.e. vitamin D metabolites), contained in the nebulized solvent, are ionized by the corona discharge needle of the interface, which applies a large voltage to the nebulized solvent/analyte mixture. The ions, i.e. precursor ions, pass through the orifice of the instrument and enter the first quadrupole. Quadrupoles 1 and 3 (Q1 and Q3) are mass filters, allowing selection of ions (i.e., "precursor" and "fragment" ions) based on their mass to charge ratio (m/z). Quadrupole 2 (Q2) is the collision cell, where ions are fragmented. The first quadrupole of the mass spectrometer (Q1) selects for molecules with the mass to charge ratios of the specific vitamin D metabolites to be analyzed. Precursor ions with the correct m/z ratios of the precursor ions of specific vitamin D metabolites are allowed to pass into the collision chamber (Q2), while unwanted ions with any other m/z collide with the sides of the quadrupole and are eliminated. Precursor ions entering Q2 collide with neutral Argon gas molecules and fragment. This process is called Collision Activated Dissociation (CAD). The fragment ions generated are passed into quadrupole 3 (Q3), where the fragment ions of the desired vitamin D metabolites are selected while other ions are eliminated.

[0048] The methods of the invention may involve MS/MS performed in either positive or negative ion mode. Using standard methods well known in the art, one of ordinary skill is capable of identifying one or more fragment ions of a particular precursor ion of a vitamin D metabolite that can be used for selection in quadrupole 3 (Q3). Preferably, at least one fragment ion of the method is specific for the particular vitamin D metabolite of which detection is desired. A specific fragment ion for a particular vitamin D metabolite is one that

will not be formed in significant amounts by other molecules with similar molecular structures. In contrast a non-specific fragment ion is one that is formed by related molecules other than the desired analyte. Therefore, detection of non-specific fragment ions alone is not reliable for distinguishing the desired vitamin D metabolite from other molecules that form the same or similar fragment ions. Specific fragment ions of a particular vitamin D metabolite can be identified by testing various molecular standards (e.g. vitamin D metabolites other than the metabolite to be detected) to determine whether fragment ions formed by the vitamin D metabolite of interest are also formed by other molecules with similar structures or features. In certain particularly preferred embodiments, a specific fragment ion is identified by testing at least one molecular standard that forms a precursor ion with the same m/z as the vitamin D metabolite to be detected.

[0049] If the precursor ion of a vitamin D metabolite of interest includes an alcohol or amine group, fragment ions are commonly formed that represent a dehydration or deamination of the precursor ion, respectively. In the case of precursor ions that include an alcohol group, such fragment ions formed by dehydration are caused by a loss of one or more water molecules from the precursor ion (i.e., where the difference in m/z between the precursor ion and fragment ion is about 18 for the loss of one water molecule, or about 36 for the loss of two water molecules, etc.). In the case of precursor ions that include an amine group, such fragment ions formed by deamination are caused by a loss of one or more ammonia molecules (i.e. where the difference in m/z between the precursor ion and fragment ion is about 17 for the loss of one ammonia molecule, or about 34 for the loss of two ammonia molecules, etc.). Likewise, precursor ions that include one or more alcohol and amine groups commonly form fragment ions that represent the loss of one or more water molecules and/or one or more ammonia molecules (e.g., where the difference in m/z between the precursor ion and fragment ion is about 35 for the loss of one water molecule and the loss of one ammonia molecule). Generally, the fragment ions that represent dehydrations or deaminations of the precursor ion are not specific fragment ions for a particular analyte. For example, MS/MS performed to detect 25OHD₂ by selecting for a precursor ion at 413 m/z (i.e. the protonated and hydrated ion) and detecting a fragment ion of 395 m/z (representing a dehydration of the 413 m/z precursor ion) would not be able to distinguish 25OHD₂ from 1 α (OH)D₂ which would form the same precursor and fragment ions. Therefore a 395 m/z fragment ion is not a specific fragment ion for 25OHD₂ or 1 α (OH)D₂. Accordingly, in preferred embodiments of

the invention, MS/MS is performed such that at least one fragment ion of a vitamin D metabolite is detected that does not represent only a loss of one or more water molecules and/or a loss of one or more ammonia molecules from the precursor ion.

[0050] As ions collide with the detector they produce a pulse of electrons that are converted to a digital signal. The acquired data is relayed to a computer, which plots counts of the ions collected versus time. The resulting mass chromatograms are similar to chromatograms generated in traditional HPLC methods. The areas under the peaks corresponding to particular ions, or the amplitude of such peaks, are measured and the area or amplitude is correlated to the amount of the analyte (vitamin D metabolite) of interest. In certain embodiments, the area under the curves, or amplitude of the peaks, for fragment ion(s) and/or precursor ions are measured to determine the amount of a vitamin D metabolite. As described above, the relative abundance of a given ion can be converted into an absolute amount of the original analyte, i.e., vitamin D metabolite, using calibration standard curves based on peaks of one or more ions of an internal molecular standard, such as $^6\text{D-25OHD}_3$.

[0051] In certain aspects of the invention, the quantity of various ions is determined by measuring the area under the curve or the amplitude of the peak and a ratio of the quantities of the ions is calculated and monitored (i.e. "daughter ion ratio monitoring"). In certain embodiments of the method, the ratio(s) of the quantity of a precursor ion and the quantity of one or more fragment ions of a vitamin D metabolite can be calculated and compared to the ratio(s) of a molecular standard of the vitamin D metabolite similarly measured. In embodiments where more than one fragment ion of a vitamin D metabolite is monitored, the ratio(s) for different fragment ions may be determined instead of, or in addition to, the ratio of the fragment ion(s) compared to the precursor ion. In embodiments where such ratios are monitored, if there is a substantial difference in an ion ratio in the sample as compared to the molecular standard, it is likely that a molecule in the sample is interfering with the results. To the contrary, if the ion ratios in the sample and the molecular standard are similar, then there is increased confidence that there is no interference. Accordingly, monitoring such ratios in the samples and comparing the ratios to those of authentic molecular standards may be used to increase the accuracy of the method.

[0052] In certain aspects of the invention, MS/MS is performed in positive ion mode with the first quadrupole (Q1) tuned to select for precursor ions with a mass charge ratio

corresponding to protonated and dehydrated ions of vitamin D metabolites. The mass/charge ratio (m/z) for the protonated and dehydrated precursor vitamin D metabolite ions is about 383.16 m/z for 25-hydroxyvitamin D₃ and about 395.30 m/z for 25-hydroxyvitamin D₂. In embodiments where the samples are spiked with hexadeuterated 25OHD₃ (⁶D-25OHD₃) for use as an internal standard, the mass/charge ratio (m/z) of the protonated and dehydrated ⁶D-25OHD₃ precursor ion is about 389.20. In certain preferred embodiments of the invention, the mass/charge ratio (m/z) for the 25-hydroxyvitamin D₃ precursor ion is about 383.16 and the m/z for at least one 25-hydroxyvitamin D₃ fragment ion is about 211.35. In related embodiments, the m/z for the 25-hydroxyvitamin D₂ precursor ion is about 395.30 and the 25-hydroxyvitamin D₂ fragment ion(s) include one or more ions selected from the group consisting of ions with mass/charge ratios (m/z) of about 179.10, about 209.20 and about 251.30. In embodiments where the samples are spiked with ⁶D-25OHD₃ for use as an internal standard the mass/charge ratio (m/z) for the protonated and dehydrated ⁶D-25OHD₃ precursor ion is about 389.20 and the fragment ion(s) may include a fragment ion with a m/z of about 211.30.

[0053] In other aspects, MS/MS is performed in positive ion mode with the first quadrupole (Q1) tuned to select for precursor ions with a mass charge ratio corresponding to protonated and hydrated ions of vitamin D metabolites. The mass/charge ratio (m/z) for the protonated and hydrated precursor vitamin D metabolite ions are about 401 m/z for 25-hydroxyvitamin D₃ and about 413 m/z for 25-hydroxyvitamin D₂. In certain preferred embodiments of the invention, the mass/charge ratio (m/z) for the 25-hydroxyvitamin D₃ precursor ion is about 401 and the 25-hydroxyvitamin D₃ fragment ion(s) include one or more ions selected from the group consisting of ions with mass/charge ratios (m/z) of about 365 and about 383 m/z ; more preferably the fragment ions include an ion with a m/z of about 365. In related embodiments, the m/z for the protonated and hydrated 25-hydroxyvitamin D₂ precursor ion is about 413 m/z , and the 25-hydroxyvitamin D₂ fragment ion(s) include one or more ions with mass/charge ratios (m/z) of about 377 and about 395; more preferably the fragment ions include an ion with a mass charge ratio of about 377.

[0054] In particularly preferred embodiments of the invention, the presence or absence or amount of two or more vitamin D metabolites in a sample are detected in a single assay using the above described MS/MS methods.

[0055] Mass spectrometry instruments can vary slightly in determining the mass of a given analyte. Thus, the term “about” in the context of mass of an ion or the m/z of an ion refers to +/- 0.5 atomic mass unit.

[0056] The following examples serve to illustrate the invention. These examples are in no way intended to limit the scope of the invention.

EXAMPLES

Example 1: Determination of 25-hydroxyvitamin D₃ and 25-hydroxyvitamin D₂ by LC-MS/MS

[0057] Using a Perkin-Elmer MultiProbe II (S/N 432400) robotic liquid handler, human serum samples were first extracted using a protein precipitation method by adding 42.5 μ l of serum to 170 μ l of methanol (1:4 ratio of serum:methanol) in a 96-well plate format. For validation-related experiments, the methanol was spiked with hexadeuterated 25OHD₃ (⁶D-25OHD₃) as an internal standard. The 96 well plates were centrifuged to remove precipitated protein, leaving the vitamin D metabolites in the supernatant. The supernatants were then transferred to an HPLC autosampler for loading to the LC-MS/MS analyzer.

[0058] LC-MS/MS was performed using a Thermo Finnigan LC-MS/MS analyzer (Thermo Finnigan Quantum TSQ (S/N: TQU00655)) with an atmospheric pressure chemical ionization (APCI) source as the detector. An autosampler was used to inject 50 μ L of extracted sample supernatant onto an HPLC column. Liquid chromatography was performed with a Cohesive Technologies Aria TX-4 (S/N: SJCTX409) LC system with Waters Symmetry C18 5 μ m 4.6x50mm columns using 100% methanol as the mobile phase. After the analytes eluted and the detector window completed acquisition, the system was washed with 85% Mobile phase A and then re-equilibrated with Mobile phase B for a run time of 5 minutes. Mobile phase A was 0.1% formic acid in HPLC-grade water and mobile phase B was 100% methanol.

[0059] The flow of liquid solvent exiting the HPLC column entered the heated nebulizer interface of the Thermo Finnigan LC-MS/MS analyzer. The solvent/analyte mixture was first converted to vapor in the heated tubing of the interface. The analytes, contained in the

nebulized solvent, were ionized (a positive charge added) by the corona discharge needle of the interface, which applies a large voltage to the nebulized solvent/analyte mixture. The ions pass through the orifice of the instrument and enter the first quadrupole. Quadrupoles 1 and 3 (Q1 and Q3) are mass filters, allowing selection of ions based on their mass to charge ratio (m/z). Quadrupole 2 (Q2) is the collision cell, where ions are fragmented.

[0060] The first quadrupole of the mass spectrometer (Q1) selected for molecules with the mass to charge ratios of (protonated and dehydrated) 25OHD₂, 25OHD₃ and ⁶D-25OHD₃. Ions with these m/z ratios (see table below) were allowed to pass into the collision chamber (Q2), while unwanted ions with any other m/z collide with the sides of the quadrupole and are eliminated. Ions entering Q2 collide with neutral Argon gas molecules and fragment. The fragment ions generated are passed into quadrupole 3 (Q3), where the fragment ions of 25OHD₂, 25OHD₃ and ⁶D-25OHD₃ were selected (see table below) and other ions are eliminated. The following mass transitions were used for detection and quantitation during validation:

Table 1. Mass transitions for selected vitamin D metabolites

| Compound | Precursor Ion (m/z) | Fragment Ions (m/z) |
|-----------------------------------|-------------------------|-------------------------|
| 25OHD ₂ | 395.30 | 179.10, 251.30, 209.20 |
| 25OHD ₃ | 383.16 | 211.35 |
| ⁶ D-25OHD ₃ | 389.20 | 211.30 |

[0061] As ions collide with the detector they produce a pulse of electrons that are converted to a digital signal. The acquired data is relayed to a computer, which plots counts of the ions collected versus time. The resulting mass chromatograms are similar to chromatograms generated in traditional HPLC methods.

[0062] Area ratios of the analyte and internal standard (Hexadeuterated 25-Hydroxyvitamin D₃, ⁶D-25OHD₃) peaks were used to construct calibration curves, which were then used to calculate analyte concentrations. Using the calibration curves, the concentrations of 25OHD₂ and 25OHD₃ were quantitated in the patient samples.

Example 2: Intra-assay and Inter-assay Precision.

[0063] Stock solutions of 25OHD₂ and 25OHD₃ were added to pooled serum to produce a Low Pool (20-25 ng/mL of each metabolite), a Medium Pool (45-55 ng/mL of each metabolite) and a High Pool (100-110 ng/mL). Pooled patient serum was used for the Medium and Low Pools, and stripped serum from Golden West Biologicals, Product #SP1070, was used for the Low Pool. Twenty aliquots from each of the Low, Medium and High Pools were analyzed in a single assay using the LC-MS/MS protocols described in Example 1. The following precision values were determined:

Table 2. Intra-Assay Variation: 25-Hydroxyvitamin D₂ (25OHD₂)

| | Low 092804-L | Medium 092804-M | High 092804-H |
|------------------------|-----------------|--------------------|------------------|
| 1 | 26.5 | 46.4 | 103.3 |
| 2 | 23.2 | 51.1 | 96.7 |
| 3 | 23.1 | 52.4 | 107.8 |
| 4 | 21.6 | 50.3 | 104.5 |
| 5 | 26.3 | 47.5 | 96.2 |
| 6 | 25.1 | 54.4 | 98.5 |
| 7 | 25.9 | 54.6 | 100.0 |
| 8 | 21.9 | 50.1 | 110.1 |
| 9 | 23.4 | 50.8 | 97.6 |
| 10 | 23.5 | 53.2 | 105.1 |
| 11 | 22.2 | 52.9 | 105.9 |
| 12 | 24.0 | 54.6 | 94.5 |
| 13 | 26.2 | 49.4 | 93.4 |
| 14 | 24.1 | 59.0 | 113.0 |
| 15 | 25.8 | 52.9 | 112.4 |
| 16 | 23.9 | 59.2 | 113.4 |
| 17 | 29.5 | 52.4 | 107.7 |
| 18 | 24.2 | 50.0 | 115.5 |
| 19 | 19.8 | 53.5 | 114.9 |
| 20 | 26.3 | 60.2 | 126.6 |
| Average (ng/mL) | 24.3 | 52.7 | 105.9 |
| Std Dev | 2.2 | 3.6 | 8.6 |
| CV (%) | 9.0 | 6.9 | 8.1 |

Table 3. Intra-Assay Variation: 25-Hydroxyvitamin D₃ (25OHD₃)

| | Low 092804-L | Medium 092804-M | High 092804-H |
|------------------------|-----------------|--------------------|------------------|
| 1 | 22.7 | 43.6 | 99.1 |
| 2 | 22.4 | 45.3 | 93.5 |
| 3 | 22.4 | 50.7 | 98.2 |
| 4 | 21.0 | 40.1 | 95.9 |
| 5 | 21.8 | 41.5 | 82.0 |
| 6 | 20.8 | 42.2 | 97.4 |
| 7 | 22.9 | 50.1 | 96.0 |
| 8 | 19.0 | 42.0 | 106.7 |
| 9 | 21.8 | 44.2 | 96.6 |
| 10 | 23.4 | 49.5 | 94.9 |
| 11 | 21.8 | 46.5 | 97.9 |
| 12 | 20.7 | 49.9 | 87.1 |
| 13 | 25.4 | 44.7 | 85.5 |
| 14 | 24.5 | 48.0 | 101.5 |
| 15 | 25.1 | 45.8 | 101.5 |
| 16 | 22.5 | 52.0 | 104.7 |
| 17 | 29.2 | 45.9 | 107.7 |
| 18 | 19.5 | 49.3 | 107.6 |
| 19 | 18.1 | 49.6 | 109.4 |
| 20 | 24.8 | 49.3 | 116.1 |
| Average (ng/mL) | 22.5 | 46.5 | 99.0 |
| Std Dev | 2.5 | 3.5 | 8.4 |
| CV (%) | 11.2 | 7.5 | 8.5 |

[0064] Each of the Low, Medium and High Pools described above were also analyzed to determine inter-assay precision. Four aliquots from each pool were analyzed over five different assays using the LC-MS/MS protocols described in Example 1. The following precision values were determined:

Table 4. Inter-Assay Variation: 25-Hydroxyvitamin D₂ (25OHD₂)

| | Low 092804-L | Medium 092804-M | High 092804-H |
|------------------------|-----------------|--------------------|------------------|
| 1 | 26.5 | 46.4 | 103.3 |
| 2 | 23.2 | 51.1 | 96.7 |
| 3 | 23.1 | 52.4 | 107.8 |
| 4 | 21.6 | 50.3 | 104.5 |
| 5 | 26.3 | 47.5 | 96.2 |
| 6 | 25.1 | 54.4 | 98.5 |
| 7 | 25.9 | 54.6 | 100.0 |
| 8 | 21.9 | 50.1 | 110.1 |
| 9 | 23.4 | 50.8 | 97.6 |
| 10 | 23.5 | 53.2 | 105.1 |
| 11 | 22.2 | 52.9 | 105.9 |
| 12 | 24.0 | 54.6 | 94.5 |
| 13 | 26.2 | 49.4 | 93.4 |
| 14 | 24.1 | 59.0 | 113.0 |
| 15 | 25.8 | 52.9 | 112.4 |
| 16 | 23.9 | 59.2 | 113.4 |
| 17 | 29.5 | 52.4 | 107.7 |
| 18 | 24.2 | 50.0 | 115.5 |
| 19 | 19.8 | 53.5 | 114.9 |
| 20 | 26.3 | 60.2 | 126.6 |
| Average (ng/mL) | 24.3 | 52.7 | 105.9 |
| Std Dev | 2.2 | 3.6 | 8.6 |
| CV (%) | 9.0 | 6.9 | 8.1 |

Table 5. Inter-Assay Variation: 25-Hydroxyvitamin D₃ (25OHD₃)

| | Low 092804-L | Medium 092804-M | High 092804-H |
|------------------------|-----------------|--------------------|------------------|
| 1 | 22.7 | 43.6 | 99.1 |
| 2 | 22.4 | 45.3 | 93.5 |
| 3 | 22.4 | 50.7 | 98.2 |
| 4 | 21.0 | 40.1 | 95.9 |
| 5 | 21.8 | 41.5 | 82.0 |
| 6 | 20.8 | 42.2 | 97.4 |
| 7 | 22.9 | 50.1 | 96.0 |
| 8 | 19.0 | 42.0 | 106.7 |
| 9 | 21.8 | 44.2 | 96.6 |
| 10 | 23.4 | 49.5 | 94.9 |
| 11 | 21.8 | 46.5 | 97.9 |
| 12 | 20.7 | 49.9 | 87.1 |
| 13 | 25.4 | 44.7 | 85.5 |
| 14 | 24.5 | 48.0 | 101.5 |
| 15 | 25.1 | 45.8 | 101.5 |
| 16 | 22.5 | 52.0 | 104.7 |
| 17 | 29.2 | 45.9 | 107.7 |
| 18 | 19.5 | 49.3 | 107.6 |
| 19 | 18.1 | 49.6 | 109.4 |
| 20 | 24.8 | 49.3 | 116.1 |
| Average (ng/mL) | 22.5 | 46.5 | 99.0 |
| Std Dev | 2.5 | 3.5 | 8.4 |
| CV (%) | 11.2 | 7.5 | 8.5 |

Example 3: Analytical Sensitivity: Limit of Detection and Limit of Quantitation Studies.

[0065] To determine the limit of detection of the assay, blank diluent was analyzed 17 times within a single run using the LC-MS/MS protocols described in Example 1. The mean and standard deviation were then calculated. The limit of detection was determined as 2 SD above the mean of the blank peak area ratio based on a back calculation of peak area ratio against the calibration curve. The limits of detection were as follows:

25OHD₂: 3.0 ng/mL
25OHD₃: 3.5 ng/mL

[0066] To determine the limit of quantitation, stock solutions of 25OHD₂ and 25OHD₃ were used to generate standard curves with the following concentrations: 0, 2, 4, 8, 16, 32, 64 and 128 ng/mL. The diluted samples of the standard curve were analyzed in quadruplicate over five assays using the LC-MS/MS assay described in Example 1. The results of the study were as follows:

Table 6. Limit of Quantitation Study Results: 25-Hydroxyvitamin D₂ (25OHD₂)

| | #1 | #2 | #3 | #4 | #5 | Summary | |
|--------------------|------|------|------|------|------|-----------------------|-------|
| 0 ng/mL | -1.2 | -1.3 | -0.5 | -1.2 | -1.1 | Average (ng/mL) | -1.0 |
| | -1.1 | -1.4 | -0.9 | -1.3 | -0.7 | Standard Deviation | 0.6 |
| | -1.1 | -1.5 | -0.8 | NA | -2.6 | C of V (%) | 59.0 |
| | 0.4 | -1.7 | -0.6 | -0.5 | -0.6 | Accuracy (%) | N/A |
| 2 ng/mL | 3.1 | 2.3 | 2.0 | 2.1 | 3.3 | Average (ng/mL) | 1.9 |
| | 2.1 | 2.7 | 2.1 | 2.4 | 1.9 | Standard Deviation | 0.6 |
| | 1.1 | 1.9 | 1.9 | 0.9 | 1.3 | C of V (%) | 33.7 |
| | 1.2 | 1.1 | 2.1 | 1.8 | 1.4 | Accuracy (%) | 103.9 |
| 4 ng/mL | 3.5 | 3.9 | 5.0 | 4.5 | 4.8 | Average (ng/mL) | 3.9 |
| | 4.0 | 3.0 | 3.8 | 3.8 | 4.7 | Standard Deviation | 0.7 |

| | | | | | | | |
|----------------------|-------|--------------|-------|-------|-------------|-----------------------|-------|
| | 3.6 | 2.9 | 2.8 | 3.1 | 2.0* | C of V (%) | 17.1 |
| | 4.1 | 4.6 | 4.5 | 4.4 | 3.7 | Accuracy (%) | 101.7 |
| 8 ng/mL | 10.2 | 9.1 | 9.1 | 8.8 | 9.8 | Average (ng/mL) | 8.6 |
| | | | | | | Standard Deviation | 0.8 |
| | 7.8 | 8.1 | 7.9 | 8.4 | 9.0 | | |
| | 8.6 | 8.3 | 7.4 | 8.4 | 7.5 | C of V (%) | 9.3 |
| | 10.2 | 8.4 | 8.0 | 8.1 | 8.6 | Accuracy (%) | 93.2 |
| 16 ng/mL | 16.0 | 14.8 | 14.4 | 16.7 | 18.3 | Average (ng/mL) | 16.0 |
| | | | | | | Standard Deviation | 1.1 |
| | 15.5 | 15.6 | 15.3 | 16.7 | 16.8 | | |
| | 16.6 | 16.7 | 16.8 | 15.8 | 15.2 | C of V (%) | 7.1 |
| | 14.1 | 17.6 | 16.1 | 16.7 | 14.1 | Accuracy (%) | 100.1 |
| 32 ng/mL | 31.3 | 39.9* | 29.9 | 33.2 | 32.7 | Average (ng/mL) | 31.8 |
| | | | | | | Standard Deviation | 1.9 |
| | 31.7 | 30.5 | 32.4 | 34.0 | 32.5 | | |
| | 29.5 | 31.2 | 30.2 | 35.7 | 28.7 | C of V (%) | 6.0 |
| | 32.9 | 34.7 | 32.4 | 30.8 | 29.2 | Accuracy (%) | 100.7 |
| 64 ng/mL | 66.5 | 62.2 | 68.5 | 62.6 | 68.8 | Average (ng/mL) | 64.6 |
| | | | | | | Standard Deviation | 3.2 |
| | 66.6 | 67.8 | 67.3 | 58.9 | 61.5 | | |
| | 64.4 | 61.4 | 63.7 | 63.5 | 61.3 | C of V (%) | 4.9 |
| | 63.7 | 60.7 | 65.4 | 70.8 | 65.9 | Accuracy (%) | 99.1 |
| 128 ng/mL | 125.1 | 128.2 | 123.4 | 127.8 | 124.1 | Average (ng/mL) | 126.9 |
| | | | | | | Standard Deviation | 3.5 |
| | 127.6 | 134.4 | 127.3 | 128.4 | 132.1 | | |
| | 128.9 | 124.5 | 128.5 | 126.5 | 131.7 | C of V (%) | 2.8 |
| | 126.1 | 119.7 | 127.9 | 121.2 | 125.0 | Accuracy (%) | 100.8 |

Table 7. Limit of Quantitation Study Results: 25-Hydroxyvitamin D₃ (25OHD₃)

| | Day #1 (11/19/04- 1) | Day #2 (11/19/04- 2) | Day #3 (11/22/04- 1) | Day #4 (11/23/04- 1) | Day #5 (11/23/04- 2) | Summary | |
|----------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|-----------------|------|
| 0 | -0.5 | -0.9 | -0.3 | 0.2 | -0.6 | Average (ng/mL) | -0.7 |

| | | | | | | | |
|-----------------|------|------|------|------|------|--------------------|-------|
| ng/mL | -0.7 | -1.3 | 0.3 | -1.1 | -1.0 | Standard Deviation | 0.6 |
| | 0.0 | -1.0 | -1.1 | -1.3 | 0.6 | C of V (%) | 86.4 |
| | -0.3 | -1.5 | -1.1 | -0.8 | -1.0 | Accuracy (%) | N/A |
| 2 ng/mL | 2.6 | 1.5 | 2.4 | 2.5 | 1.7 | Average (ng/mL) | 1.9 |
| | 1.7 | 0.9 | 3.2 | 2.2 | 1.9 | Standard Deviation | 0.6 |
| | 1.8 | 1.8 | 2.0 | 1.1 | 2.2 | C of V (%) | 31.4 |
| | 1.3 | 2.4 | 1.5 | 1.1 | 2.5 | Accuracy (%) | 104.7 |
| 4 ng/mL | 3.9 | 3.5 | 4.8 | 4.0 | 3.4 | Average (ng/mL) | 3.8 |
| | 3.0 | 4.1 | 3.1 | 3.9 | 3.8 | Standard Deviation | 0.8 |
| | 4.5 | 3.7 | 2.9 | 3.5 | 2.4 | C of V (%) | 19.7 |
| | 4.0 | 5.2 | 3.8 | 3.8 | 5.5 | Accuracy (%) | 104.1 |
| 8 ng/mL | 10.3 | 9.3 | 7.2 | 8.4 | 8.6 | Average (ng/mL) | 8.7 |
| | 10.6 | 8.5 | 7.2 | 10.0 | 9.6 | Standard Deviation | 1.1 |
| | 7.4 | 10.3 | 9.0 | 9.4 | 8.4 | C of V (%) | 13.1 |
| | 9.3 | 7.7 | 8.5 | 7.6 | 6.8 | Accuracy (%) | 91.9 |
| 16 ng/mL | 15.9 | 15.6 | 16.2 | 18.6 | 17.1 | Average (ng/mL) | 16.0 |
| | 13.8 | 16.3 | 14.0 | 17.2 | 15.3 | Standard Deviation | 1.4 |
| | 15.6 | 16.1 | 15.1 | 18.8 | 16.1 | C of V (%) | 8.5 |
| | 14.8 | 17.2 | 17.0 | 14.3 | 15.3 | Accuracy (%) | 99.9 |
| 32 ng/mL | 31.1 | 35.8 | 29.6 | 32.8 | 28.0 | Average (ng/mL) | 31.7 |
| | 30.8 | 29.9 | 31.8 | 33.0 | 32.8 | Standard Deviation | 1.9 |
| | 31.6 | 30.9 | 29.5 | 35.7 | 31.2 | C of V (%) | 6.2 |
| | 31.0 | 34.2 | 31.8 | 30.6 | 31.7 | Accuracy (%) | 101.0 |
| 64 ng/mL | 65.9 | 64.6 | 64.4 | 64.8 | 67.8 | Average (ng/mL) | 64.8 |
| | 67.4 | 62.9 | 62.4 | 60.7 | 57.2 | Standard Deviation | 3.1 |
| | 68.9 | 64.2 | 62.1 | 64.0 | 64.7 | C of V (%) | 4.8 |

| | | | | | | | |
|--------------|-------|-------|-------|-------|-------|-----------------------|-------|
| | 63.2 | 64.2 | 66.8 | 67.7 | 71.1 | Accuracy (%) | 98.8 |
| 128 ng/mL | 128.9 | 124.5 | 126.4 | 125.3 | 122.8 | Average (ng/mL) | 127.1 |
| | 129.7 | 135.7 | 128.3 | 125.9 | 128.7 | Standard Deviation | 4.7 |
| | 125.5 | 123.3 | 127.2 | 127.7 | 135.4 | C of V (%) | 3.7 |
| | 121.3 | 121.8 | 137.8 | 121.4 | 124.1 | Accuracy (%) | 100.7 |

Example 4: Assay Reportable Range and Linearity.

[0067] To establish the linearity of the vitamin D metabolite LC-MS/MS assay, a MultiProbe automated liquid handler robot independently constructed two standard curves by serially diluting a stock solution containing 128 ng/mL 25OHD₂ and 128 ng/mL 25OHD₃ in 25OHD with a solution of 5% Bovine Serum Albumin Fraction V dissolved in 0.01M PBS. The standard curve samples were analyzed using the LC-MS/MS protocols described in Example 1. This process routinely produced standard curves with R² values of 0.99 or higher for each analyte for the range of 4 – 128 ng/mL.

[0068] To determine whether patient samples can also be diluted in a linear fashion, a total of eight samples were serially diluted with 25OHD diluent. Two samples were patient pools (Medium and High Control Pools), three were patient samples with high 25OHD₂ values and three were patient samples with high 25OHD₃ values. All samples were analyzed using the LC-MS/MS protocols described in Example 1. As shown in figure 1 and figure 2, each sample diluted in a linear fashion (R² > 0.98), demonstrating the linear range of the assay.

[0069] Additionally, solutions of 25OHD₂ and 25HOD₃ at 512 ng/mL in 5% Bovine Serum Albumin Fraction V dissolved in 0.01M PBS were prepared and then serially diluted to 8 ng/mL. Each sample was extracted and run in duplicate using the LC-MS/MS protocols described in Example 1. As shown in figure 3, each of these curves was linear (R² > 0.99).

Example 5: Accuracy of LC-MS/MS Vitamin D Assay.

[0070] The stock solutions of 25OHD₂ and 25OHD₃ were quantified based upon the absorbance of the concentrated (10-50 µg/mL) stock solutions in the ultraviolet spectrum. The *cis*-triene chromophore present in all vitamin D compounds has a peak absorbance of 264 nm, which is dependent upon the analyte concentration. The molar extinction coefficient

of $18.3 \text{ mM}^{-1} \text{ cm}^{-1}$ was determined using purified, dessicated ergocalciferol and cholecalciferol and was used to determine the concentration of stock solutions for 25OHD_2 and 25OHD_3 .

[0071] To determine the ability to recover vitamin D metabolites from spiked serum samples, three patient pools with known levels of 25OHD_2 , 25OHD_3 were spiked with two levels of 25OHD_2 , 25OHD_3 and both 25OHD_2 and 25OHD_3 together. Each sample was extracted and run in duplicate using the LC-MS/MS protocols described in Example 1. The recovery was calculated by dividing the expected result by the observed result.

Table 8. Recovery of 25-hydroxylated vitamin D metabolites from spiked samples.

| | 25OHD ₂ (ng/mL) | 25OHD ₃ (ng/mL) | 25OHD ₂ (% Recovery) | 25OHD ₃ (% Recovery) |
|-------------------------------------|-------------------------------|-------------------------------|------------------------------------|------------------------------------|
| Pool#1 | 58 | 51 | - | - |
| Pool#1+ 20 ng/mL 25OHD ₂ | 75 | 52 | 104 | - |
| Pool#1+ 20 ng/mL 25OHD ₃ | 52 | 68 | - | 105 |
| Pool#1+ 20 ng/mL of both | 72 | 68 | 109 | 105 |
| Pool#1+ 50 ng/mL 25OHD ₂ | 104 | 50 | 105 | - |
| Pool#1+ 50 ng/mL 25OHD ₃ | 56 | 109 | - | 93 |
| Pool#1+ 50 ng/mL of both | 104 | 110 | 104 | 92 |

| | 25OHD ₂ (ng/mL) | 25OHD ₃ (ng/mL) | 25OHD ₂ (% Recovery) | 25OHD ₃ (% Recovery) |
|-------------------------------------|-------------------------------|-------------------------------|------------------------------------|------------------------------------|
| Pool#2 | 53 | 47 | - | - |
| Pool#2+ 20 ng/mL 25OHD ₂ | 76 | 51 | 97 | - |
| Pool#2+ 20 ng/mL 25OHD ₃ | 52 | 66 | - | 101 |
| Pool#2+ 20 ng/mL of both | 70 | 65 | 105 | 103 |
| Pool#2+ 50 ng/mL 25OHD ₂ | 107 | 53 | 96 | - |
| Pool#2+ 50 ng/mL 25OHD ₃ | 57 | 109 | - | 88 |
| Pool#2+ 50 ng/mL of both | 100 | 105 | 103 | 92 |

| | 25OHD ₂ (ng/mL) | 25OHD ₃ (ng/mL) | 25OHD ₂ (% Recovery) | 25OHD ₃ (% Recovery) |
|-------------------------------------|-------------------------------|-------------------------------|------------------------------------|------------------------------------|
| Pool#3 | 53 | 47 | - | - |
| Pool#3+ 20 ng/mL 25OHD ₂ | 69 | 44 | 106 | - |
| Pool#3+ 20 ng/mL 25OHD ₃ | 55 | 75 | - | 90 |
| Pool#3+ 20 ng/mL of both | 74 | 69 | 98 | 97 |
| Pool#3+ 50 ng/mL 25OHD ₂ | 96 | 48 | 107 | - |
| Pool#3+ 50 ng/mL 25OHD ₃ | 53 | 114 | - | 85 |
| Pool#3+ 50 ng/mL of both | 105 | 113 | 98 | 85 |

Example 6: Comparison LC-MS/MS vitamin D metabolite assay and RIA procedures.

[0072] A total of 1,057 patient samples were assayed using the LC-MS/MS methods described in Example 1 and a vitamin D radioimmunoassay commercially available from DiaSorin. Figure 4 shows the correlation between detection of total 25-hydroxyvitamin D using the LC-MS/MS assay and a commercially available radioimmunoassay kit; the R² value was 0.5082 with a slope of 0.9684.

Example 7: Selectivity of the LC-MS/MS assay.

[0073] Samples containing various commercially available vitamin D metabolites and analogues at a concentration of 100 ng/mL were prepared. The samples were extracted and run in duplicate using the LC-MS/MS methods described in Example 1. None of the tested compounds exhibited detectable cross reactivity in the 25OHD₂ and 25OHD₃ assays.

Table 9. Cross-Reactivity of the LC-MS/MS method with various vitamin D analogues and metabolites.

| | Compound Mass (Da) | Cross-Reactivity (25OHD₂) | Cross-Reactivity (25OHD₃) |
|---|---------------------------|---|---|
| 25-Hydroxyvitamin D ₂ (25OHD ₂) | 412 | (100%) | ND |
| 25-Hydroxyvitamin D ₃ (25OHD ₃) | 400 | ND | (100%) |
| Internal Standard (⁶ D-25OHD ₃) | 406 | ND | ND |
| Vitamin D ₂ (Ergocalciferol) | 396 | ND | ND |
| Vitamin D ₃ (Cholecalciferol) | 384 | ND | ND |
| 1 α ,25(OH) ₂ D ₂ | 428 | ND | ND |
| 1 α ,25(OH) ₂ D ₃ | 416 | ND | ND |
| 25,26(OH) ₂ D ₃ | 416 | ND | ND |
| 1 α (OH)D ₂ (Doxercalciferol) | 412 | ND | ND |
| 1 α (OH)D ₃ (Alfacalcidol) | 400 | ND | ND |

[0074] The contents of the articles, patents, and patent applications, and all other documents and electronically available information mentioned or cited herein, are hereby incorporated by reference in their entirety to the same extent as if each individual publication was specifically and individually indicated to be incorporated by reference. Applicants reserve the right to physically incorporate into this application any and all materials and information from any such articles, patents, patent applications, or other physical and electronic documents.

[0075] The inventions illustratively described herein may suitably be practiced in the absence of any element or elements, limitation or limitations, not specifically disclosed herein. Thus, for example, the terms “comprising”, “including,” containing”, etc. shall be read expansively and without limitation. Additionally, the terms and expressions employed herein have been used as terms of description and not of limitation, and there is no intention in the use of such terms and expressions of excluding any equivalents of the features shown and described or portions thereof, but it is recognized that various modifications are possible within the scope of the invention claimed. Thus, it should be understood that although the present invention has been specifically disclosed by preferred embodiments and optional features, modification and variation of the inventions embodied therein herein disclosed may be resorted to by those skilled in the art, and that such modifications and variations are considered to be within the scope of this invention.

[0076] The invention has been described broadly and generically herein. Each of the narrower species and subgeneric groupings falling within the generic disclosure also form part of the invention. This includes the generic description of the invention with a proviso or negative limitation removing any subject matter from the genus, regardless of whether or not the excised material is specifically recited herein.

[0077] Other embodiments are within the following claims. In addition, where features or aspects of the invention are described in terms of Markush groups, those skilled in the art will recognize that the invention is also thereby described in terms of any individual member or subgroup of members of the Markush group.

THAT WHICH IS CLAIMED IS:

1. A method for determining the presence or amount of a vitamin D metabolite in a sample by tandem mass spectrometry, comprising:
 - (a) generating a protonated and dehydrated precursor ion of said vitamin D metabolite;
 - (b) generating one or more fragment ions of said precursor ion; and
 - (c) detecting the presence or amount of one or more of said ions generated in step (a) or (b) or both and relating the detected ions to the presence or amount of said vitamin D metabolite in said sample.
2. The method of claim 1, wherein said sample is subjected to a purification step prior to ionization.
3. The method of claim 2, wherein said purification step comprises chromatography.
4. The method of claim 3, wherein said chromatography comprises high performance liquid chromatography (HPLC).
5. The method of claim 2, wherein said purification step comprises protein precipitation.
6. The method of claim 5, wherein said purification step does not include high turbulence liquid chromatography (HTLC).
7. The method of claim 2, wherein said purification step comprises chiral chromatography.
8. The method of claim 1, wherein said sample or said vitamin D metabolite is not subjected to gas-chromatography prior to said ionization step.
9. The method of claim 1, wherein said vitamin D metabolite is one or more vitamin D metabolites selected from the group consisting of 25-hydroxyvitamin D₃, 25-hydroxyvitamin D₂, 1,25-dihydroxyvitamin D₃ and 1,25-dihydroxyvitamin D₂.
10. The method of claim 1, wherein said vitamin D metabolite is 25-hydroxyvitamin D₃.

11. The method of claim 10, wherein said precursor ion has a mass/charge ratio (m/z) of about 383.16.
12. The method of claim 11, wherein said precursor ion has a mass/charge ratio (m/z) of about 383.16, and said one or more fragment ions comprise a fragment ion with an m/z of about 211.35.
13. The method of claim 1, wherein said vitamin D metabolite is 25-hydroxyvitamin D₂.
14. The method of claim 13, wherein said precursor ion has a mass/charge ratio (m/z) of about 395.30.
15. The method of claim 13, wherein said precursor ion has a mass/charge ratio (m/z) of about 395.30, and said one or more fragment ions comprise one or more ions selected from the group consisting of ions with a m/z of about 179.10, about 209.20, and about 251.30.
16. The method of claim 15, wherein said one or more fragment ions comprise two or more ions selected from the group consisting of ions with a m/z of about 179.10, about 209.20, and about 251.30.
17. The method of claim 1, wherein said vitamin D metabolite is 1,25-dihydroxyvitamin D₂.
18. A method for determining the presence or amount of two or more vitamin D metabolites in a sample in a single assay, said method comprising:
 - (a) ionizing said two or more vitamin D metabolites, if present in said sample, to generate protonated and dehydrated precursor ions specific for each of said two or more vitamin D metabolites;
 - (b) generating one or more fragment ions of each of said precursor ions; and
 - (c) detecting the presence or amount of one or more of said ions generated in step (a) or (b) or both and relating the detected ions to the presence or amount of said two or more vitamin D metabolites in said sample.
19. The method of claim 18, wherein said sample subjected to a purification step prior to said ionization step.

20. The method of claim 19, wherein said purification step comprises chromatography.
21. The method of claim 20, wherein said chromatography comprises is high performance liquid chromatography (HPLC).
22. The method of claim 19, wherein said purification step comprises protein precipitation and does not include high turbulence liquid chromatography.
23. The method of claim 18, wherein said sample or said two or more vitamin D metabolites are not subjected to gas-chromatography prior to said ionization step.
24. The method of claim 18, wherein said vitamin D metabolites comprise 25-hydroxyvitamin D₃ and 25-hydroxyvitamin D₂.
25. The method of claim 24, wherein the precursor ion of 25-hydroxyvitamin D₃ has a mass/charge ratio (m/z) of about 383.16 and the precursor ion of 25-hydroxyvitamin D₂ has a mass/charge ratio (m/z) of about 395.30.
26. The method of claim 25, wherein the one or more fragment ions for 25-hydroxyvitamin D₃ comprise an ion with an m/z of about 383.16.
27. The method of claim 25, wherein the one or more fragment ions for 25-hydroxyvitamin D₂ comprise one or more ions selected from the group consisting of ions with a m/z of about 179.10, about 209.20, and about 251.30.
28. The method of claim 18, wherein said two or more vitamin D metabolites comprise one or more vitamin D metabolites selected from the group consisting of 25-hydroxyvitamin D₃, 25-hydroxyvitamin D₂, 1,25-dihydroxyvitamin D₃ and 1,25-dihydroxyvitamin D₂.
29. The method of claim 18, wherein said two or more vitamin D metabolites comprise 1,25-dihydroxyvitamin D₃ and 1,25-dihydroxyvitamin D₂.

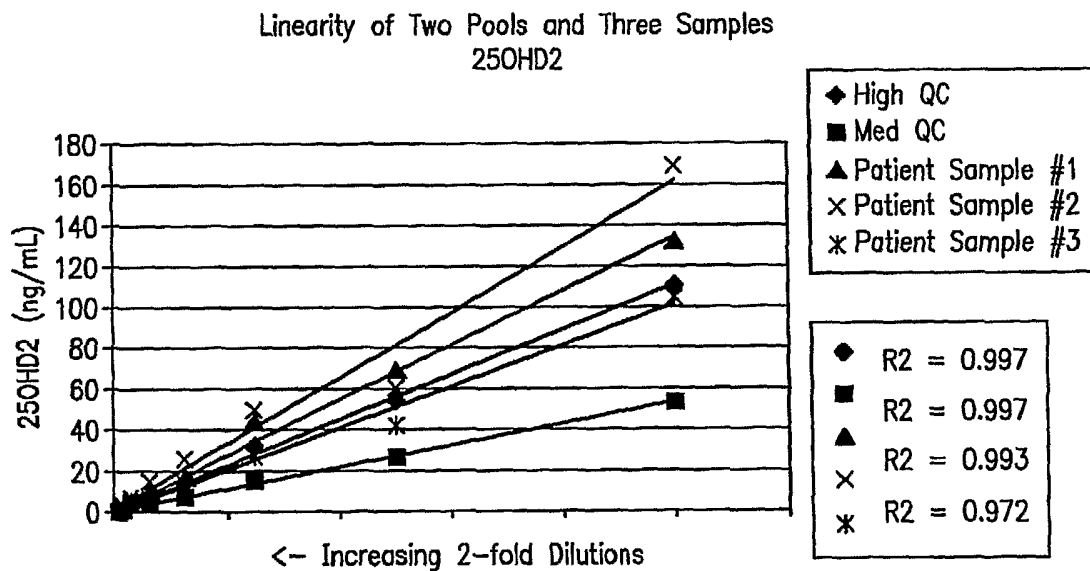


FIG. 1

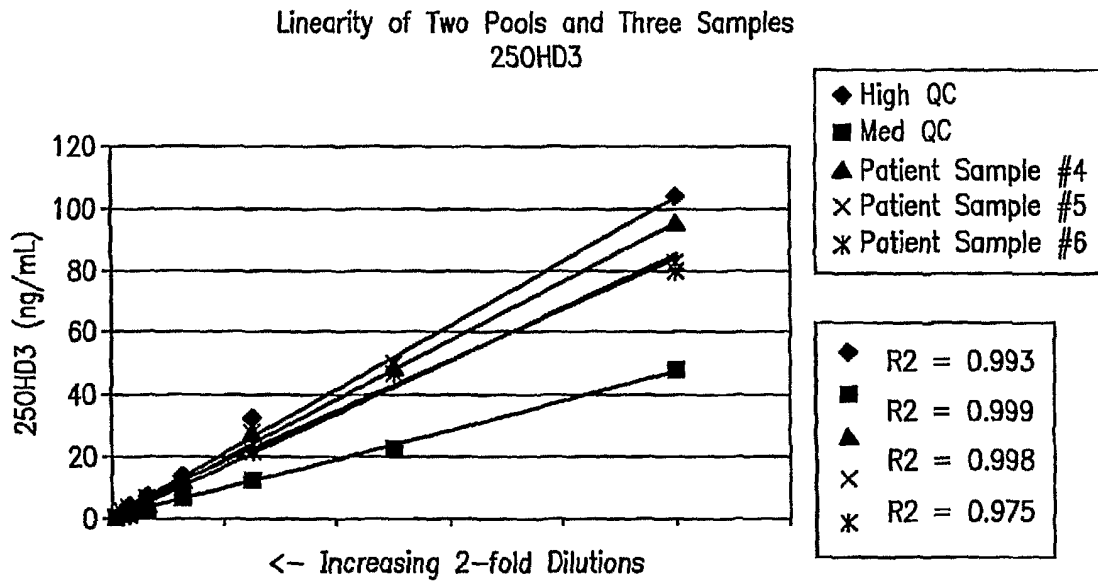


FIG. 2

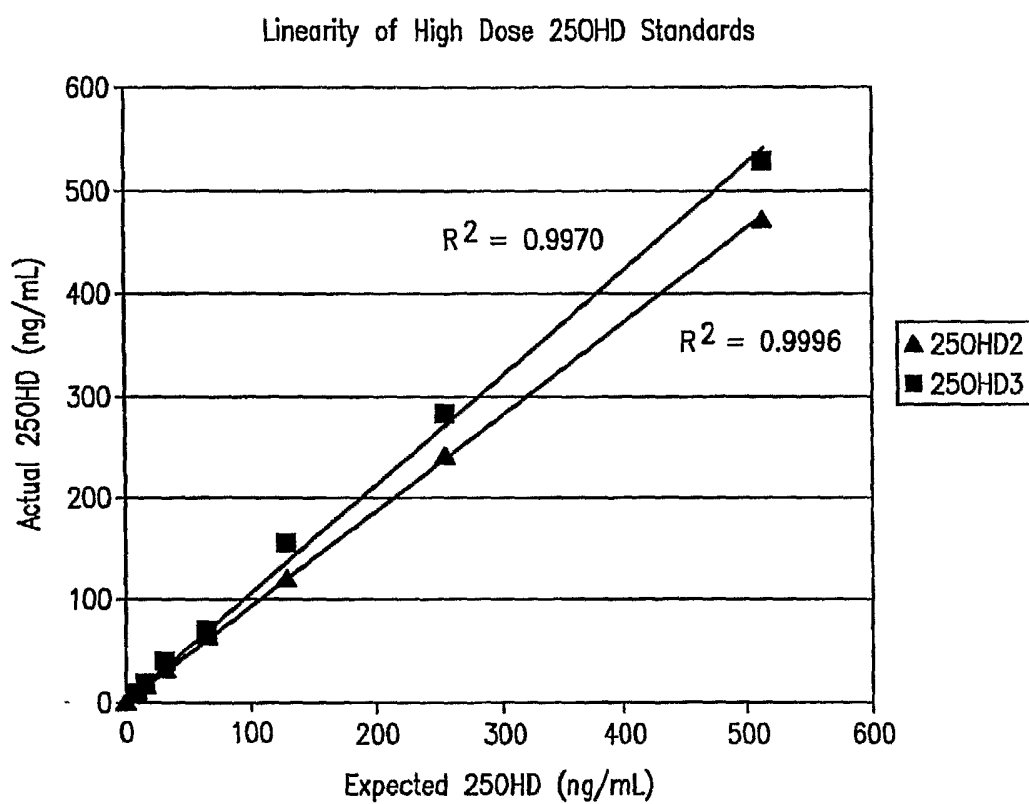


FIG. 3

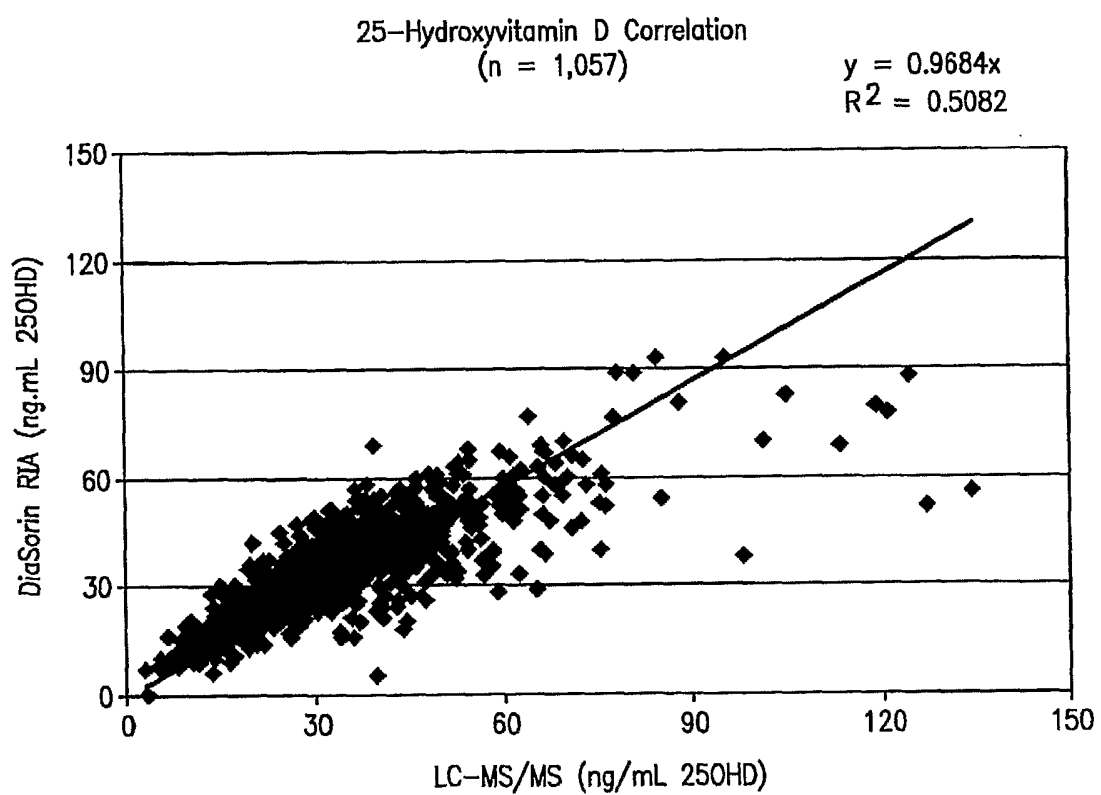


FIG. 4