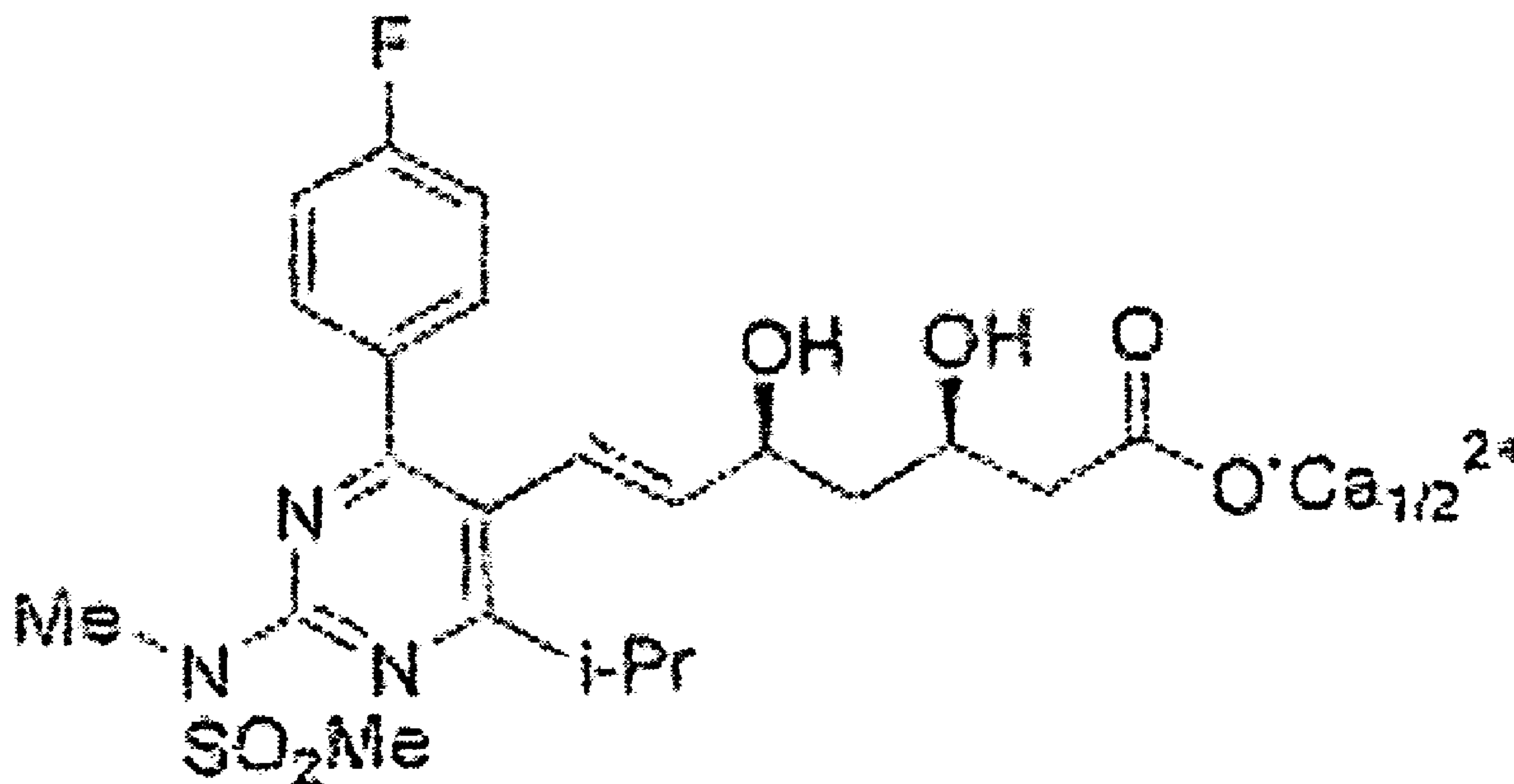




(86) **Date de dépôt PCT/PCT Filing Date:** 2006/06/22  
(87) **Date publication PCT/PCT Publication Date:** 2006/12/28  
(45) **Date de délivrance/Issue Date:** 2015/05/05  
(85) **Entrée phase nationale/National Entry:** 2007/12/12  
(86) **N° demande PCT/PCT Application No.:** EP 2006/006008  
(87) **N° publication PCT/PCT Publication No.:** 2006/136408  
(30) **Priorité/Priority:** 2005/06/24 (SI P200500187)

(51) **Cl.Int./Int.Cl. C07D 239/42** (2006.01),  
**A61K 31/505** (2006.01)  
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(54) **Titre : PROCÉDE DE PREPARATION DE ROSUVASTATINE CALCIQUE PURE AMORPHE**  
(54) **Title: PROCESS FOR PREPARING PURE AMORPHOUS ROSUVASTATIN CALCIUM**



**Formula 1**

(57) **Abrégé/Abstract:**

A new process for preparing pure amorphous rosuvastatin calcium, substantially free of impurities, is disclosed. A process comprising hydrolysing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin, preferably methyl rosuvastatin or tert-butyl rosuvastatin, with a base, e.g. sodium hydroxide, in the presence of an aprotic solvent, preferably tetrahydrofuran and N,N-dimethyl acetamide, or in the presence of a mixture of an aprotic solvent and water, to obtain a solution of rosuvastatin salt, which may be converted to another



**(57) Abrégé(suite)/Abstract(continued):**

rosuvastatin salt using another cation, e.g. with calcium cation to obtain rosuvastatin calcium. Rosuvastatin amine salts may be obtained as well. In another preferred aspect of the invention rosuvastatin free acid may be converted to various rosuvastatin salts, e.g. to rosuvastatin calcium, rosuvastatin sodium or various rosuvastatin amine salts, including rosuvastatin solvates, e.g. rosuvastatin calcium hydrate. Rosuvastatin calcium is useful in the treatment of hyperlipidemia, hypercholesterolemia and atherosclerosis.

## (12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization  
International Bureau



PCT



(43) International Publication Date  
28 December 2006 (28.12.2006)

(10) International Publication Number  
**WO 2006/136408 A3**

(51) International Patent Classification:

C07D 239/42 (2006.01) A61K 31/505 (2006.01)

(21) International Application Number:

PCT/EP2006/006008

(22) International Filing Date:

22 June 2006 (22.06.2006)

(25) Filing Language:

English

(26) Publication Language:

English

(30) Priority Data:

P200500187 24 June 2005 (24.06.2005) SI

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(81) Designated States (unless otherwise indicated, for every kind of national protection available): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BW, BY, BZ, CA, CH, CN,

CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, HN, HR, HU, ID, IL, IN, IS, JP, KE, KG, KM, KN, KP, KR, KZ, LA, LC, LK, LR, LS, LT, LU, LV, LY, MA, MD, MG, MK, MN, MW, MX, MZ, NA, NG, NI, NO, NZ, OM, PG, PH, PL, PT, RO, RS, RU, SC, SD, SE, SG, SK, SL, SM, SY, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, ZA, ZM, ZW.

(84) Designated States (unless otherwise indicated, for every kind of regional protection available): ARIPO (BW, GH, GM, KE, LS, MW, MZ, NA, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IS, IT, LT, LU, LV, MC, NL, PL, PT, RO, SE, SI, SK, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

Declaration under Rule 4.17:

— as to applicant's entitlement to apply for and be granted a patent (Rule 4.17(ii))

Published:

— with international search report

(88) Date of publication of the international search report:

19 April 2007

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

(54) Title: PROCESS FOR PREPARING PURE AMORPHOUS ROSUVASTATIN CALCIUM

(57) Abstract: A new process for preparing pure amorphous rosuvastatin calcium, substantially free of impurities, is disclosed. A process comprising hydrolysing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin, preferably methyl rosuvastatin or tert-butyl rosuvastatin, with a base, e.g. sodium hydroxide, in the presence of an aprotic solvent, preferably tetrahydrofuran and N,N-dimethyl acetamide, or in the presence of a mixture of an aprotic solvent and water, to obtain a solution of rosuvastatin salt, which may be converted to another rosuvastatin salt using another cation, e.g. with calcium cation to obtain rosuvastatin calcium. Rosuvastatin amine salts may be obtained as well. In another preferred aspect of the invention rosuvastatin free acid may be converted to various rosuvastatin salts, e.g. to rosuvastatin calcium, rosuvastatin sodium or various rosuvastatin amine salts, including rosuvastatin solvates, e.g. rosuvastatin calcium hydrate. Rosuvastatin calcium is useful in the treatment of hyperlipidemia, hypercholesterolemia and atherosclerosis.

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## **Process for preparing pure amorphous rosuvastatin calcium**

### **Field of the invention**

The present invention relates to a new process for preparing pure amorphous rosuvastatin calcium, substantially free of impurities.

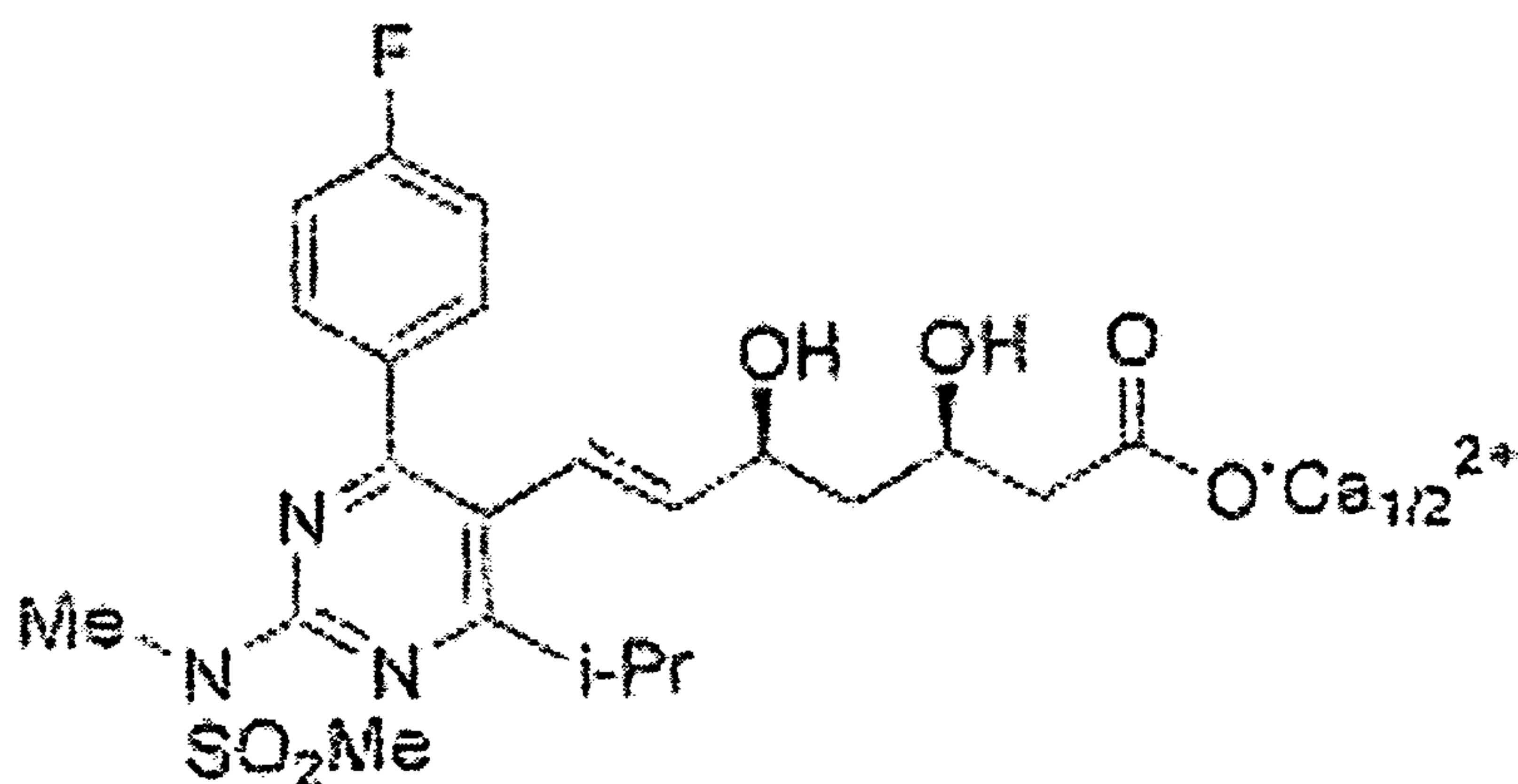
### **Background of the invention**

Rosuvastatin is generic name for (+)-7-[4-(4-fluorophenyl)-6-isopropyl-2-(N-methyl-N-methylsulfonylamino)pyrimidin-5-yl]-(3R,5S)-dihydroxy-(E)-heptenoic acid administered in the therapy as its calcium salt as commercial drug, and illustrated in Formula 1 hereinafter, which compound is an inhibitor of the enzyme 3-hydroxy-3-methylglutaryl-coenzyme A reductase (HMG CoA reductase), useful in the treatment of hyperlipidemia, hypercholesterolemia and atherosclerosis. Rosuvastatin and the synthesis of rosuvastatin calcium was first disclosed in patent EP-B-521 471; in the last two steps of the whole synthesis provided by hydrolysis of methyl ester of rosuvastatin (methyl rosuvastatin) in polar solvent, e.g. ethanol, in the presence of a base, following by isolation of sodium salt of rosuvastatin (sodium rosuvastatinate) and converting said sodium salt of rosuvastatin with a water soluble calcium salt under aqueous conditions to calcium salt of rosuvastatin.

WO 2005/023778 discloses a process for the preparation of rosuvastatin calcium by conversion of C<sub>1</sub> to C<sub>4</sub> alkyl ester of rosuvastatin, preferably *tert*-butyl ester of rosuvastatin with a base, preferably sodium hydroxide, in the presence of a C<sub>1</sub> to C<sub>4</sub> alcohol, preferably ethanol, to a solution of rosuvastatin salt, e.g. its sodium salt and converted said salt into rosuvastatin calcium by adding a source of calcium to said solution.

A novel crystalline form of rosuvastatin calcium can be prepared by crystallization of amorphous form of rosuvastatin calcium from a mixture of: (i) water and acetonitrile in the ratio of 1 : 1 by volume; (ii) water and acetone in the ratio of 1 : 1 by volume; or water, methanol and *tert*-butyl methyl ether in the ratio of 1 : 1 : 1 by volume, what is described in WO 2000/042024.

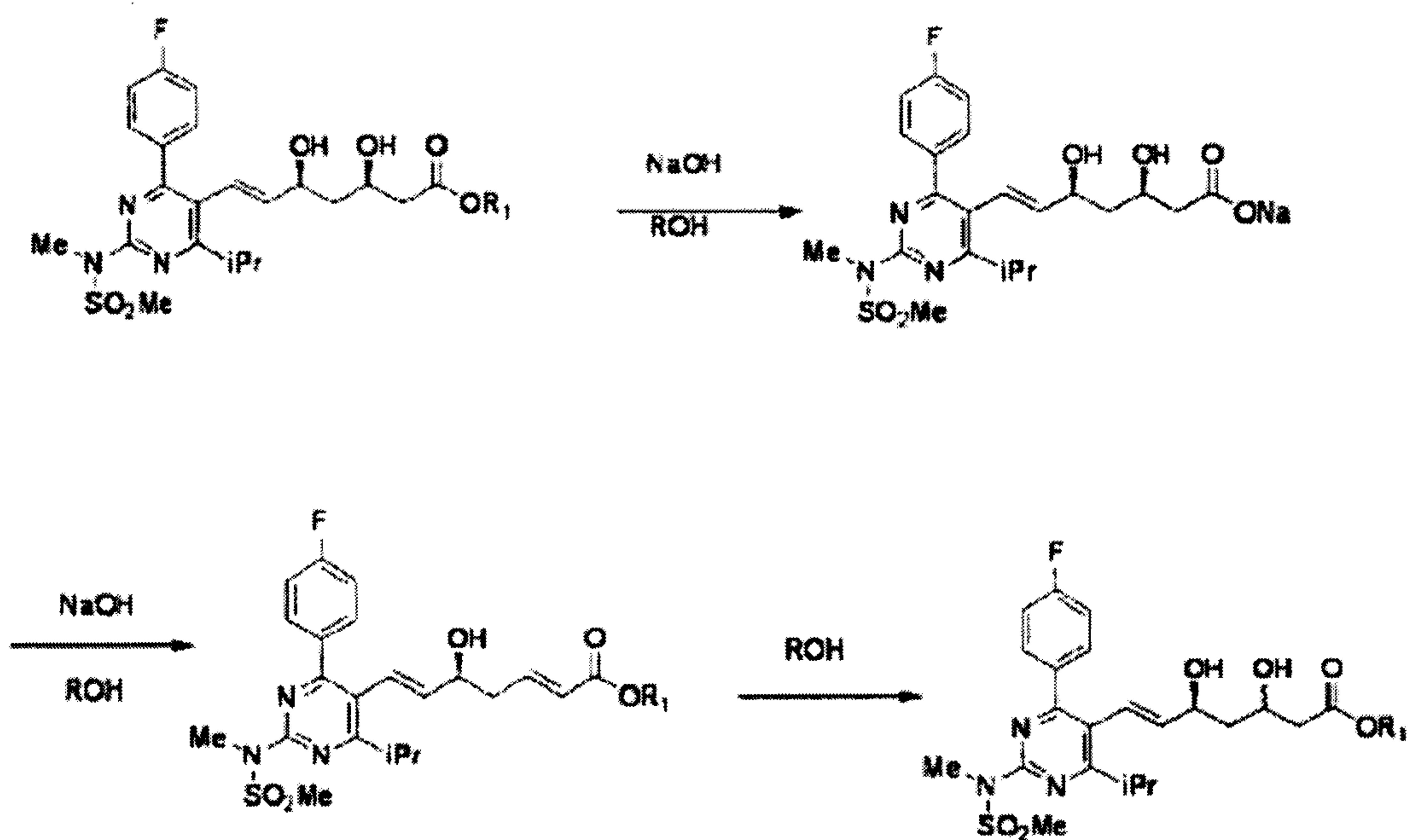
WO 2005/040134 describes amorphous rosuvastatin calcium having a purity of more than 99% obtained from known crystalline rosuvastatin calcium. Said crystalline form may be prepared by crystallization of amorphous rosuvastatin calcium which process is known from the prior art.



Formula 1

It is well known that alkali metal salts of organic carboxylic acids are often hygroscopic what may cause problems at isolation. Indeed the isolation of sodium salt of rosuvastatin, which can be an intermediate in preparing rosuvastatine calcium salt, might be unrepeatable in yield and physical state what depends on the reaction conditions and evaporation of solvents, which is difficult to control. International publication WO 05/23778 tried to avoid said problems without isolating rosuvastatin sodium salt by extraction of impurities from its aqueous solution, but by using C<sub>1</sub> to C<sub>4</sub> alcohols as reaction medium a risk of conversion into specific impurities still existed. Namely, it is known that  $\beta$ -hydroxy acids in alcoholic alkali solution are

submitted to dehydration what may lead after realkoxylation into special side products (see Scheme 1, wherein R and R<sub>1</sub> independently denotes C<sub>1</sub> to C<sub>5</sub> alkyl group), O-alkyl rosuvastatin, e.g. O-ethyl rosuvastatin.



Scheme 1

Figure 1 shows results of stress stability test, from which is obvious that rosuvastatin was submitted to specific degradation in ethanol as alcoholic medium.

The specific impurity, designated as NN 20-47.143 is O-ethyl ether derivative (see Scheme 1 wherein R denotes ethyl (Et) group), was detected in 0.35% area by HPLC in comparison with acetone solution in which no such impurity was observed.

Therefore a need for an efficient process for preparing pure rosuvastatin calcium, without any significant amounts of side products, still exists.

### **Summary of the invention**

In a first aspect the present invention provides a process for producing calcium salt of rosuvastatin comprising:



## 3a

- a) hydrolysing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with a base in the presence of an aprotic solvent or with a base in the presence of a mixture of aprotic solvent and water, to obtain a solution of rosuvastatin salt,
- b) converting thus obtained rosuvastatin salt with a source of calcium to obtain rosuvastatin calcium
- c) isolating the calcium salt of rosuvastatin.

In another aspect of the invention the C<sub>1</sub> to C<sub>5</sub> ester of rosuvastatin may be converted to solid sodium salt of rosuvastatine, which is isolated from the obtained solution; said isolated salt having excellent handling physical properties.

In a further aspect of the invention the rosuvastatin free acid ("rosuvastatinic acid") may be isolated according to the procedure comprising:

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- a) hydrolysing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with a base in an aprotic solvent, optionally diluted with water,
- b) removing excess of organic aprotic solvent,
- c) optionally diluting water poor solution with additional water,
- d) optionally washing thus obtained aqueous solution of rosuvastatin salt of a base with a water immiscible organic solvent,
- e) acidifying aqueous solution of rosuvastatin salt of a base,
- f) extracting the resulted rosuvastatin acid with the water immiscible organic solvent,
- g) removing the water immiscible organic solvent from the obtained extract to obtain isolated rosuvastatin acid.

Yet another aspect of the invention presents a conversion of thus obtained rosuvastatin acid (rosuvastatinic acid) to any salt of rosuvastatin, e.g. to its sodium or calcium salt, by simple adding cation source in a solvent to rosuvastatin acid.

By conversion of rosuvastatin acid into its calcium salt in nonaqueous medium substantially anhydrous rosuvastatin calcium salt may be obtained.

Yet another aspect of the invention presents a new more convenient procedure of preparing of monohydrates of rosuvastatin calcium.

Rosuvastatin calcium prepared by such improved procedures according to the invention has at least 99.5 % of chromatographic purity; moreover when using very pure starting C<sub>1</sub> to C<sub>5</sub> rosuvastatin ester of the invention more than 99.8 % purity, even more, in some cases more than 99.9 % of chromatographic purity of desired rosuvastatin calcium may be obtained.

The term "chromatographic purity" means purity as determined by HPLC ("High Pressure Liquid Chromatography").



## 4a

In another aspect, the present invention provides a process for producing amorphous rosuvastatin calcium which comprises:

- a) hydrolysing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with:
  - (i) sodium hydroxide in the presence of tetrahydrofuran, or
  - (ii) sodium hydroxide in the presence of a mixture of tetrahydrofuran and water,to obtain a solution of rosuvastatin sodium;
- b) converting thus obtained rosuvastatin sodium with a source of calcium to obtain rosuvastatin calcium;
- c) redigesting the obtained rosuvastatin calcium in water;
- d) stirring the obtained rosuvastatin calcium; and
- e) isolating amorphous rosuvastatin calcium.

In another aspect, the present invention provides a process for producing amorphous rosuvastatin calcium having a residual sodium cation content below 0.1% by weight, which comprises:

- a) hydrolysing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with:
  - (i) sodium hydroxide in the presence of tetrahydrofuran, or
  - (ii) sodium hydroxide in the presence of a mixture of tetrahydrofuran and water,to obtain a solution of rosuvastatin sodium;
- b) converting thus obtained rosuvastatin sodium with a source of calcium to obtain rosuvastatin calcium;
- c) redigesting the obtained rosuvastatin calcium in water;
- d) stirring the obtained rosuvastatin calcium with a dispersion making device with high speed rotation unit; and
- e) isolating amorphous rosuvastatin calcium having a residual sodium cation content below 0.1% by weight.

In another aspect, the present invention provides a process for producing amorphous rosuvastatin calcium, comprising:

## 4b

- a) hydrolyzing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with sodium hydroxide or potassium hydroxide in an amide solvent in the absence of water, optionally in combination with ether;
- b) removing said amide solvent to obtain a solid residue of rosuvastatin sodium or rosuvastatin potassium;
- c) optionally washing the obtained solid residue of rosuvastatin sodium or rosuvastatin potassium with an ether and isolating solid rosuvastatin sodium or rosuvastatin potassium;
- d) combining the solid residue of rosuvastatin sodium or rosuvastatin potassium with water to obtain an aqueous solution of rosuvastatin sodium or rosuvastatin potassium;
- e) adding a source of calcium to the resulted solution to precipitate rosuvastatin calcium; and
- f) isolating amorphous rosuvastatin calcium salt.

In another aspect, the present invention provides a process for producing amorphous rosuvastatin calcium having a residual sodium cation content below 0.1% by weight, which comprises:

- a) hydrolyzing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with sodium hydroxide in an amide solvent in the absence of water, optionally in combination with ether;
- b) removing said amide solvent to obtain a solid residue of rosuvastatin sodium;
- c) optionally washing the obtained solid residue of rosuvastatin sodium with an ether and isolating solid rosuvastatin sodium;
- d) combining the solid residue of rosuvastatin sodium with water to obtain an aqueous solution of rosuvastatin sodium;
- e) adding a source of calcium to the resulted solution to precipitate rosuvastatin calcium;
- f) redigesting the obtained rosuvastatin calcium in water,
- g) stirring the obtained rosuvastatin calcium with a dispersion making device with high speed rotation unit; and
- h) isolating amorphous rosuvastatin calcium salt having a residual sodium cation content below 0.1% by weight.

## 4c

In another aspect, the present invention provides a process for preparing crystalline rosuvastatin calcium monohydrate, comprising contacting the amorphous rosuvastatin calcium prepared according to the above-mentioned process with a mixture of acetone and water.

In another aspect, the present invention provides a process for producing anhydrous amorphous rosuvastatin calcium salt, comprising contacting rosuvastatin free acid with calcium 2-ethylhexanoate in iso-butyl acetate.

In another aspect, the present invention provides amorphous rosuvastatin calcium having residual amounts of sodium cation below 0.1 % by weight.

In another aspect, the present invention provides a pharmaceutical formulation comprising rosuvastatin calcium with residual amounts of sodium cation, wherein said rosuvastatin calcium is amorphous rosuvastatin calcium having sodium cation content below 0.1 % by weight

In another aspect, the present invention provides the use of rosuvastatin calcium prepared by the above-mentioned process for the treatment and/or prophylaxis of hyperlipidemia, hypercholesterolemia and atherosclerosis.

In another aspect, the present invention provides the use of rosuvastatin calcium prepared by the above-mentioned process for the preparation of a medicament for the treatment and/or prophylaxis of hyperlipidemia, hypercholesterolemia and atherosclerosis.

In another aspect, the present invention provides a process for preparing a pharmaceutical formulation comprising rosuvastatin calcium, said process comprising preparing rosuvastatin calcium by the above described process, and formulating said rosuvastatin calcium into a pharmaceutical formulation comprising rosuvastatin calcium.



4d

In another aspect, the present invention provides a pharmaceutical formulation comprising rosuvastatin calcium prepared according to above described process and

a method of treatment of hyperlipidemia, hypercholesterolemia and atherosclerosis, comprising the step of administering the said pharmaceutical formulation to the mammal in need thereof.

The starting ester may be preferably methyl ester of rosuvastatin, more preferably *tert*-butyl ester of rosuvastatin (*tert*-butyl rosuvastatin).

Any aprotic solvent of step a) of above described process may be used in the reaction, preferably tetrahydrofuran (THF) or *N,N*-dimethylacetamide (DMA).

The base used in the above described process of step a) is selected from the group consisting of sodium hydroxide, potassium hydroxide, and their analogues.

Any appropriate source of calcium source may be used, preferably calcium chloride or calcium acetate.

The water immiscible solvent used in above steps d) and f) is selected from the group consisting of C<sub>1</sub>-C<sub>4</sub> esters, e.g. acetate esters, preferably ethyl acetate (AcOEt), ethers, chlorinated hydrocarbons, cyclic hydrocarbons.

As acid for acidifying aqueous solution in step e) of above described process any inorganic acid or organic acid may be used, such as hydrochloric acid or sulfuric acid.

The term "rosuvastatin acid" means (+)-7-[4-(4-fluorophenyl)-6-isopropyl-2-(*N*-methyl-*N*-methylsulfonylamino)pyrimidin-5-yl]-(3*R*,5*S*)-dihydroxy-(*E*)-6-heptenoic acid.

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**Brief description of the drawings**

Figure 1 shows comparison of results of stress stability tests, in which rosuvastatin was submitted to specific degradation in ethanol solution and acetone solution, respectively.

Figure 2 shows results of X-Ray powder analysis – diffraction angles (2 theta) of crystalline rosuvastatin calcium monohydrate salt as described in Example 7.

**Detailed description of the invention**

An object of the present invention is to find a novel process for the preparation of pure amorphous rosuvastatin calcium, substantially free of impurities, which would avoid the use of alcohols, e.g. C<sub>1</sub> to C<sub>4</sub> alcohol as a reaction medium known in the



prior art processes, thus resulting to elimination O-alkyl rosuvastatin impurities (see Scheme 1), e.g. O-ethyl rosuvastatin, containing in rosuvastatin calcium, prepared according to the prior art processes.

Further is the object of the present invention to find a novel process which would enable easy and simple preparation and isolation of intermediary rosuvastatin salt, e.g. rosuvastatin sodium or rosuvastatin acid, in good quality, crystallinity and repeatedly, which were difficult to isolate in the prior art processes and which would enable simple and easy conversion of said intermediary compounds to desired commercial rosuvastatin calcium.

The term "substantially free of impurities" means less than 0.5% of total impurities as measured by area percentage HPLC, preferably less than 0.2% of total impurities as measured by area percentage HPLC, most preferably less than 0.1% of total impurities as measured by area percentage HPLC.

We have unexpectedly and surprisingly found that above problem has been solved by hydrolysis of starting C<sub>1</sub> to C<sub>5</sub> alkyl esters of rosuvastatin, where instead of using C<sub>1</sub> to C<sub>4</sub> alcohols as solvent medium known in the prior art processes said hydrolysis take place in an aprotic solvents with added water. Any suitable aprotic solvent may be used, but preferably tetrahydrofuran (THF) and *N,N*-dimethylacetamide (DMA).

In the first aspect of the invention a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin, where alkyl denotes methyl, ethyl, n-propyl, *iso*-propyl, n-butyl, *iso*-butyl, *tert*-butyl, amyl and *tert*-amyl group, most preferably *tert*-butyl ester of rosuvastatin, is reacted with a base in an aprotic solvent - water mixture, preferably with sodium hydroxide as a base, in tetrahydrofuran (THF) containing medium, which may be in the mixture with other solvents, preferably THF and water mixture in all ratios, more preferably the mixture of tetrahydrofuran and water in the ratio of 5:1 to 1:5 by volume, in a further embodiment in the ratio for 4:1 to 1:3 by volume. Pure THF may be employed as medium as well. After completing of said hydrolysis reaction preferably rosuvastatin

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sodium is obtained. Hydrolysis of starting C<sub>1</sub> to C<sub>5</sub> ester of rosuvastatin may be performed at temperatures from 20°C to 60°C. The hydrolysis is accomplished in 30 minutes (at 60°C) to 2 hours (at 20°C). Thereafter the obtained

reaction mixture may be diluted with another solvent, preferably with water, if aqueous poor hydrolysis medium has been previously used. The aprotic organic solvent is then optionally removed, the residual aqueous solution is washed by water immiscible organic solvent selected from the group consisting of esters, ethers, chlorinated hydrocarbons or cyclic hydrocarbons, preferably more user friendly solvents, e.g. acetate esters, more preferably ethyl acetate (AcOEt), or cyclic hydrocarbons, to obtain aqueous solution of pure rosuvastatin alkaline salt, preferably sodium salt of rosuvastatin.

Alternatively, after said hydrolysis in water rich media, the reaction mixture is washed by water immiscible organic solvents selected from the group consisting of esters, ethers, chlorinated hydrocarbons or cyclic hydrocarbons, preferably selected from more user friendly solvents, e.g. acetate esters or cyclic hydrocarbons, more preferably ethyl acetate, without previous partial evaporation of organic component of the mixture with water, for example tetrahydrofuran.

Thereafter a source of calcium ions is added to the solution of rosuvastatin salt, preferably to the solution of rosuvastatin sodium salt, for instance calcium halogenide, preferably calcium chloride, and another calcium source, for example calcium nitrate or calcium hydroxide, calcium salt of C<sub>1</sub>-C<sub>20</sub> alkanolic acid, preferably calcium palmitate, calcium pivalate or calcium acetate. The most preferable calcium source is calcium acetate monohydrate (Ca(OAc)<sub>2</sub>·H<sub>2</sub>O). Addition of Ca(OAc)<sub>2</sub>·H<sub>2</sub>O may be performed at temperatures between 0 °C to 40 °C. Ca(OAc)<sub>2</sub>·H<sub>2</sub>O may be added in one portion or dropwise in 5 to 60 minutes. After the addition of Ca(OAc)<sub>2</sub>·H<sub>2</sub>O resulted suspension may be stirred at temperatures from 0°C to 40 °C from 30 minutes to 2 hours. The desired rosuvastatin calcium may be formed as a solid precipitate, which is conveniently isolated.

Obtained amorphous rosuvastatin calcium of said process is chromatographically pure (over 99.5 % area, mainly over 99.8 % area as measured by HPLC), but may contain various amounts of residual cations, in the case of sodium hydroxide used in said hydrolysis residual amounts of sodium cation. In order to make a conversion process more repeatable in removing most of sodium cation, special care for stirring



should be performed. So the precipitation of rosuvastatin calcium salt from the reaction mixture may be executed by stirring said reaction mixture with ultraturrax (Ultra-Turrax® is brand name of IKA Werke GmbH & Co., Staufen, Germany) for dispersion making device with high speed rotation unit) in operation speed from 6000 rpm to 30000 rpm, preferably from 15000 rpm to 25000 rpm, most preferably 16000 rpm to 20000 rpm. The obtained rosuvastatin calcium may also be redigested in fresh water and executing Ultraturax® stirring at the same conditions. In such preferred case content of sodium cation falls in desired rosuvastatin calcium below 0.1 % by weight.

Washing of cationic impurities from the desired rosuvastatin calcium product may be carried out not only in water but also in appropriate solvent / water mixture, preferably tetrahydrofuran / water mixture.

Applying the treatment of rosuvastatin calcium with the mixture of acetone and water a monohydrate of rosuvastatin calcium salt may be prepared. Obtained crystal form of said product is identical to rosuvastatin calcium monohydrate from the prior art as proved by X-ray diffraction pattern (XRD) analysis (see Figure 2). Rosuvastatin calcium monohydrate may be prepared by retreating any form of rosuvastatin calcium in the mixture of acetone and water or prepared by converting other rosuvastatin salts or rosuvastatin acid with calcium source in the mixture of acetone and water.

In the second aspect of the invention of using aprotic solvents and a base in hydrolysis of starting C<sub>1</sub> to C<sub>5</sub> rosuvastatin esters but in the absence of water the aprotic solvents are evaporated under the reduced pressure at temperatures from 10 °C to 50 °C. The obtained salt of rosuvastatin with the base used for hydrolysis may be isolated. The preferred solvent used for said anhydrous cleavage of starting C<sub>1</sub> to C<sub>5</sub> rosuvastatin esters are amides, preferably *N,N*-dimethylacetamide (DMA), optionally in a combination with another solvent selected from ethers, e.g. diethyl ether. Thus the starting C<sub>1</sub> to C<sub>5</sub> ester of rosuvastatin is hydrolyzed with a base in pure DMA or in a mixture of DMA and ether in the ratio of 1:1 to 3:1 by volume. Hydrolysis may be performed at temperatures from 0 °C to 30 °C and is

accomplished in 30 to 90 minutes. Precipitation of the desired rosuvastatin salt may be achieved with additional amounts of ethers, such as diethyl ether ( $\text{Et}_2\text{O}$ ), *tert*-butyl methyl ether ( $^t\text{BuMeO}$ ), diisopropyl ether ( $^i\text{Pr}_2\text{O}$ ). In such case after removing of solvents the resulted rosuvastatin salt, preferably rosuvastatin sodium salt is washed and isolated as easy filtrable solid. The term "reduced pressure" generally refers to a pressure of about 10 mbar to about 50 mbar.

Thereafter rosuvastatin sodium salt is dissolved in water and then a source of calcium ions is added. Before adding a source of calcium the aqueous solution of rosuvastatin salt of a base may be optionally washed with a water immiscible organic solvent followed by removing of said solvent. For instance calcium halogenide, preferably calcium chloride, another calcium source, for example calcium nitrate or calcium hydroxide, calcium salt of  $\text{C}_1\text{-C}_{20}$  alkanolic acid, preferably calcium palmitate, calcium 2-ethylhexanoate, calcium pivalate or calcium acetate may be used in said conversion. Addition of  $\text{Ca}(\text{OAc})_2 \times \text{H}_2\text{O}$  may be performed at temperatures from 0 °C to 40 °C;  $\text{Ca}(\text{OAc})_2 \times \text{H}_2\text{O}$  may be added in one portion or dropwise in 5 to 60 minutes. After the addition of  $\text{Ca}(\text{OAc})_2 \times \text{H}_2\text{O}$  suspension may be stirred at temperatures between 0 °C to 40 °C from 30 minutes to 2 hours. The obtained amorphous rosuvastatin calcium is formed as a solid precipitate, which is conveniently isolated. Also in this aspect residual amounts of sodium cation containing in rosuvastatin calcium may be reduced with special care on stirring and redigesting, as described above.

Rosuvastatin calcium salt isolated according to any of previously described procedures is in amorphous form. The final product rosuvastatin calcium may be conveniently isolated also as a solvate, preferably as rosuvastatin calcium monohydrate, which may be isolated from the mixture of acetone and water added to rosuvastatin calcium, preferably from the mixture of acetone and water in the ratio of 2:1 by volume, what is more simple procedure as already known long drawn procedure from the prior art.

Described procedure provides a desired rosuvastatin calcium of high purity where according to the first aspect of the invention the purity of so produced rosuvastatin



calcium exceeds 99.8% area, preferably 99.9% area, while according to the second aspect the produced rosuvastatin calcium exceeds 99.5 % area, preferably 99.7 % area as measured by HPLC.

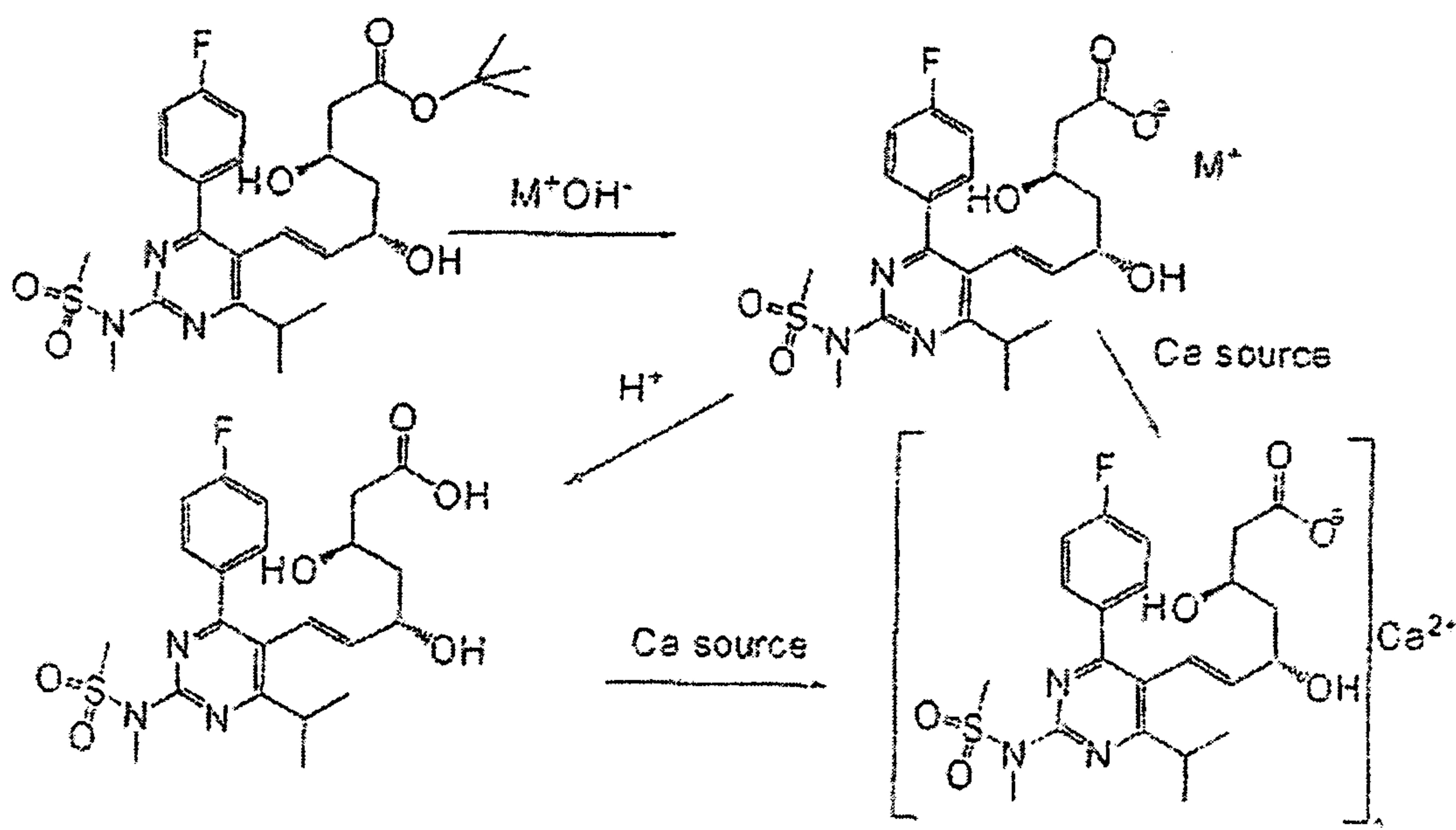
In the third aspect of the invention the use of aprotic solvents in the hydrolysis of starting C<sub>1</sub> to C<sub>5</sub> rosuvastatin esters enables an isolation of pure solid rosuvastatin acid. Performing hydrolytic conditions previously described as aprotic solvent - water mixture and a base and after removing most of organic solvents from the resulted aqueous solution of rosuvastatin salt is acidified with an inorganic acid, preferably hydrochloric acid, and rosuvastatin free acid is extracted into a water immiscible or partially miscible organic solvents, for example ethers, esters, chlorinated hydrocarbons, preferably C<sub>1</sub>-C<sub>4</sub> acetate esters, most preferably ethyl acetate (AcOEt).

Isolation of rosuvastatin solid free acid or preparation of pure rosuvastatin free acid solution according to the invention makes possible a preparation of various new amine salts of rosuvastatin and various anhydrous rosuvastatin salts. In hereinafter presented examples *tert*-octylammonium (2,4,4-trimethyl-2-pentylammonium) rosuvastatin salt and anhydrous rosuvastatin calcium salt may be prepared. The preparation of rosuvastatin amine salt may be performed in acetonitrile (MeCN) as a solvent by adding *tert*-octylamine (2,4,4-trimethyl-2-pentylamine) to the rosuvastatin free acid solution, following filtering off the rosuvastatin *tert*-amine salt precipitate after completing of the reaction.

The preparation of anhydrous amorphous rosuvastatin calcium salt may also be performed in *iso*-butyl acetate solvent by adding calcium 2-ethylhexanoate to the rosuvastatin free acid solution, following filtering off the precipitate of amorphous rosuvastatin calcium. Substantially anhydrous rosuvastatin calcium may be used in special pharmaceutical formulations of the present invention in which the absence of water is desired. The term "substantially anhydrous" means that the content of water in anhydrous rosuvastatin calcium is less than 0.1% by weight.

The various aspects of the invention are presented in the following scheme 2 and described in detail in following examples:





( $M^+$  denotes alkali metal cation, e.g. sodium cation)

Scheme 2

### Analytical method

Chromatographic purity is determined by HPLC method by the following method:

Equipment: Waters Alliance 2695 separations module, detector PDA 2996, software Empower 5.0; column: Xterra RP 18, 3  $\mu m$ , 150 x 4,6 mm; mobile phase: A: Buffer 10mM  $KH_2PO_4$ , pH=2,5; B: acetonitrile ; temperature: 45  $^{\circ}C$ ; flow rate: 1,2 ml/min; wavelength: 224 and 242 nm; injection volume: 20  $\mu l$ ; gradient table:

| t  | % A | % B |
|----|-----|-----|
| 0  | 60  | 40  |
| 15 | 60  | 40  |
| 25 | 50  | 50  |
| 25 | 30  | 70  |
| 40 | 30  | 70  |
| 41 | 60  | 0   |

Powder X-ray diffraction spectra of the sample is recorded on Siemens D-5000, CuK $\alpha$  radiation, range from 2° to 37° 2 $\theta$ , step 0.04° 2 $\theta$ , integration time 1 sec, slit V20 in 0.6.

### Example 1

#### Preparation of rosuvastatin calcium salt via rosuvastatin sodium salt solution

Rosuvastatin *tert*-butyl ester (60.0 g, 111.6 mmol) is dissolved in 500 mL of a 4:1 mixture of THF / water. The clear solution is warmed to 30 °C and 8.0 M NaOH (15 mL, 120.0 mmol) is added portionwise. The reaction mixture is stirred at 30 °C for 2 hours giving a clear yellow solution. Then THF is removed completely under the reduced pressure (20 mbar) at 40 °C. The remaining aqueous solution is diluted with water to 500 mL and washed with AcOEt (2×200 mL). After separation from the organic layer aqueous phase is distilled under the reduced pressure (20 mbar) at 40 °C to completely remove the dissolved AcOEt. The remaining clear solution of sodium rosuvastatinate (440 mL) is diluted with water (60 mL) to 500 mL and warmed to 40 °C. To a vigorously stirring solution of sodium rosuvastatinate is added dropwise Ca(OAc)<sub>2</sub>·H<sub>2</sub>O (14.8 g, 84.0 mmol in 60 mL of water) over 5 minutes at 40 °C to precipitate rosuvastatin calcium. After the complete addition the suspension is stirred further for 30 minutes at 40 °C. The white precipitate is filtered off. Then a wet white solid is suspended in water (200 mL) and vigorously stirred for 1 hour at 20 °C. The undissolved precipitate is collected by filtration, washed with water (200 mL) and dried in vacuum at 40 °C to give 48.5 g (86.8 %) of rosuvastatin calcium salt as white powder (HPLC: 99.87 %).

### Example 2

Rosuvastatin *tert*-butyl ester (60.0 g, 111.6 mmol) is dissolved in 120 mL of tetrahydrofuran (THF) and 300 ml of water treated 8.0 M NaOH (21 mL) is added portionwise. The reaction mixture is stirred at 50°C for 2 hours. Reaction mixture is allowed to cool to room temperature and washed with 2 x 540 ml of methylcyclohexane.



Aqueous phase is evaporated at 60°C under reduced pressure to 220 ml of total volume to eliminate organic solvents. The residue obtained is rediluted with degassed demi-water to 440 ml of total volume. To the resulting solution 1.0 g of charcoal is added and the suspension is stirred half an hour. Charcoal is filtered off. One half of the volume of the filtrate (220 ml of total 440 ml) 25.5 ml aqueous solution of calcium chloride (prepared from 10.5 ml 4M calcium chloride and 15 ml demi-water) is added during stirring on ice-bath. The suspension formed is treated vigorously with ultraturrax at cca 18000 rpm for 3 minutes. The precipitate is filtered off, suspended anew in 100 ml demi water and treated again with Ultraturax® at 18000 rpm for 3 minutes on ice-bath. The product is separated by filtration, washed with 30 ml ice-cold degassed demi-water, collected from the filter and dried 12 hours at 50°C in vacuum desiccator.

Yield: 25.05 g of amorphous rosuvastatin calcium (99.75% area, HPLC, 0.085 % Na)

The second aliquot of 220 ml of filtrate is treated on the same way except mechanical stirring instead of ultraturax mixing is performed. Yield 25.11 g (99.72% area, HPLC, 1.55 % Na)

### Example 3

Rosuvastatin *tert*-butyl ester (10.0 g) is dissolved in 20 mL of tetrahydrofuran and 50 ml of water treated 8.0 M NaOH (3.51 mL) is added portionwise. The reaction mixture is stirred at 50°C for 1 hours. One third (26 ml of the total 78 ml) of the resulting solution is washed with 35 ml methylcyclohexane. Methylcyclohexane phase is extracted with 3 ml demi water. Combined aqueous phases are washed with 20 ml *iso*-propyl acetate. Aqueous phase is then concentrated by evaporation under reduced pressure at 50°C to 15 - 20 ml of total volume. It is cooled on ice-bath and gradually 1.1 ml aqueous solution of 4M calcium chloride is added within a minute during stirring. It is stirred additional 30 minutes on ice-bath and the precipitated product is separated by filtration. The precipitate is washed with 4.0 ml demi water,

collected from the filter and dried at room temperature 12 hours in vacuum desiccator.

Yield: 2.22 g of amorphous rosuvastatin calcium.

Two further aliquots are washed with 20 ml ethyl acetate or 20 ml *tert*-butyl methyl ether respectively with similar yield and quality.

#### Example 4

##### Preparation of rosuvastatin sodium salt

Rosuvastatin *tert*-butyl ester (3.0 g, 5.6 mmol) is dissolved in 25 mL of a 4:1 mixture of THF / water. The clear solution is warmed to 30 °C and 8.0 M NaOH (0.75 mL, 6.0 mmol) was added portionwise. The reaction mixture is stirred at 30 °C for 2 hours giving a clear yellow solution. Then THF is removed completely under the reduced pressure (20 mbar) at 40 °C. The remaining aqueous solution is diluted with water to 25 mL and washed with AcOEt (2×10 mL). After separation from the organic layer aqueous phase is distilled under the reduced pressure (20 mbar) at 40 °C to completely remove the dissolved AcOEt. The remaining clear solution of sodium rosuvastatinate is diluted with water to 25 mL and lyophilized to afford 2.81 g (100 %) of rosuvastatin sodium salt as white powder.

#### Example 5

Rosuvastatin *tert*-butyl ester (3.0 g, 5.6 mmol) is dissolved in *N,N*-dimethylacetamide (4 mL) at ambient temperature giving a clear yellow solution. Then NaOH 8.0 M (0.75 mL, 6.0 mmol) is added dropwise to a stirred solution at ambient temperature. The reaction mixture is stirred at ambient temperature for 30 minutes giving a clear yellow solution. Then Et<sub>2</sub>O (50 mL) is added portionwise to the vigorously stirred reaction mixture which becomes immediately turbid and light yellow oil separated from the mixture. After 20 minutes of vigorous stirring oil solidified and fine white precipitate is



formed. The suspension is stirred at ambient temperature for 2.5 hours. The white precipitate (sodium rosuvastatin) is filtered off and washed with Et<sub>2</sub>O (2×20 mL).

#### Example 6

##### Preparation of rosuvastatin calcium salt from isolated rosuvastatin sodium salt

Rosuvastatin sodium as white solid from Example 5 is dissolved in water (30 mL) and solution is filtered to give clear colourless solution. The pH of solution is adjusted to 7 by addition of 1.0 M HCl (0.5 mL). To a vigorously stirring solution of sodium rosuvastatin is added dropwise Ca(OAc)<sub>2</sub>×H<sub>2</sub>O (0.74 g, 4.2 mmol in 3 mL of water) over 1 minute at ambient temperature to precipitate rosuvastatin calcium. After the complete addition the suspension is stirred further for 45 minutes at ambient temperature. The white precipitate is filtered off, washed with water (2×10 mL) and dried in vacuum at 40 °C to give 2.43 g (87.0 %) of rosuvastatin calcium salt as white powder (HPLC: 99.72 %).

#### Example 7

##### Preparation of crystalline rosuvastatin calcium monohydrate salt

To a rosuvastatin calcium salt (2.0 g, 2.0 mmol) is added 10 mL of a 2:1 mixture of acetone / water. The mixture is stirred at ambient temperature for 30 minutes to give a white turbid solution. Then additional 1 mL of a 2:1 mixture of acetone / water is added into the mixture. After 10 minutes of stirring at ambient temperature white solid precipitates abundantly from the mixture. The precipitate is collected by filtration. After drying in vacuum at 40°C 1.67 g (82.0 %) of crystalline rosuvastatin calcium monohydrate salt is obtained. Crystalline form is confirmed by comparing diffractogram (Figure 2) with reference picture from the prior art.

## Example 8

Preparation of solid rosuvastatin free acid

Rosuvastatin *tert*-butyl ester (27.0 g, 50.2 mmol) is dissolved in 225 mL of a 4:1 mixture of THF/water. The clear solution is warmed to 30 °C and 8.0 M NaOH (6.75 mL, 54.0 mmol) is added portionwise. The reaction mixture is stirred at 30 °C for 2 hours giving a clear yellow solution. Then THF is removed completely under the reduced pressure (20 mbar) at 40 °C. The remaining aqueous solution is diluted with water to 225 mL and washed with AcOEt (3×90 mL). To a vigorously stirring solution of sodium rosuvastatinate is added dropwise HCl 37 % (4.2 mL, 50.2 mmol) at ambient temperature. The obtained white emulsion of a free acid is extracted with AcOEt (150 mL). After separation from the organic layer aqueous phase is additionally extracted with AcOEt (2×50 mL). Organic layers are combined and washed with water (3×30 mL). Then AcOEt is removed under reduced pressure (20 mbar) at 40 °C. The residue is dissolved in a minimal amount of acetonitrile (MeCN) and the solvent is rapidly evaporated under reduced pressure (20 mbar) at 40 °C to give 25.48 g of the solid residue of free rosuvastatin acid.

## Example 9

Preparation of crystalline *tert*-octylammonium salt of rosuvastatin

Solid free rosuvastatin acid is then dissolved in acetonitrile (MeCN) (100 mL) to give a clear solution. To a vigorously stirring solution of a rosuvastatin free acid is added dropwise *tert*-octylamine (6.83 g, 50.2 mmol) over 1 minute at ambient temperature. In less than 10 minutes white solid precipitates abundantly from the solution, which cause solidification of the mixture. This solid is then treated with 75 mL of a 1:2 mixture of hexane / MeCN to give a dense suspension. The white precipitate is filtered and dried in vacuum at 40 °C to give 27.6 g of a white powder. This powder is suspended in hexane (100 mL) and vigorously stirred for 1 hour at ambient temperature. The undissolved precipitate is collected by filtration, washed with

hexane (50 mL) and dried in vacuum at 40 °C to give 27.4 g (89.4 %) of rosuvastatin *tert*-octylammonium salt as white crystalline powder.

#### Example 10

20.0 g of rosuvastatin *tert*-butyl ester is dissolved in the mixture of 40 ml of tetrahydrofuran and 100 ml of water, then 7.0 ml 8M NaOH is added and stirred at 40°C for 1 hour. The resulting solution is washed with 200 ml methylcyclohexane. Methylcyclohexane phase is extracted with 5 ml demi water. Aqueous phases are combined getting 140 ml of total volume of rosuvastatin sodium salt solution.

To this solution 180 ml *iso*-butyl acetate and 5.4 ml 85% orto-phosphoric acid in 21 ml of demi water is added, wherein released free rosuvastatinic acid is extracted in the organic phase and the layers formed are separated. To the organic phase 3.0 g charcoal and 30 g anhydrous magnesium sulfate are added and the resulting suspension is stirred 45 minutes. The following filtration yielded 240 ml of filtrate.

To the filtrate the solution of 84 ml of 0.4 M calcium 2-ethylhexanoate in *iso*-butyl acetate is added. Then, to the reaction mixture 240 ml of n-heptane is gradually added during stirring on the ice-bath. The solid precipitate is filtered off and washed with 120 ml *iso*-butyl acetate/n-heptane mixture (1 : 1).

Yield: 19.7 g of amorphous rosuvastatin calcium is collected after drying.



## Claims

1. A process for producing amorphous rosuvastatin calcium having a residual sodium cation content below 0.1% by weight, which comprises:
  - a) hydrolysing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with:
    - (i) sodium hydroxide in the presence of tetrahydrofuran, or
    - (ii) sodium hydroxide in the presence of a mixture of tetrahydrofuran and water, to obtain a solution of rosuvastatin sodium;
  - b) converting thus obtained rosuvastatin sodium with a source of calcium to obtain rosuvastatin calcium;
  - c) redigesting the obtained rosuvastatin calcium in water;
  - d) stirring the obtained rosuvastatin calcium with a dispersion making device with high speed rotation unit; and
  - e) isolating amorphous rosuvastatin calcium having a residual sodium cation content below 0.1% by weight.
2. The process according to claim 1, wherein the C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin is methyl rosuvastatin or *tert*-butyl rosuvastatin.
3. The process according to claim 2, wherein the C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin is *tert*-butyl rosuvastatin.
4. The process according to claim 1, wherein the ratio of tetrahydrofuran and water is from 5:1 to 1:5 by volume.
5. The process according claim 4, wherein the ratio of tetrahydrofuran and water is from 4:1 to 1:3 by volume.
6. The process according to any claim 1, wherein the source of calcium is calcium chloride, calcium nitrate, calcium hydroxide or a calcium salt of a C<sub>1</sub> to C<sub>20</sub> alkanolic acid.
7. The process according to claim 6, wherein the calcium salt of a C<sub>1</sub> to C<sub>20</sub> alkanolic acid is calcium palmitate, calcium pivalate or calcium acetate.

8. A process for producing amorphous rosuvastatin calcium having a residual sodium cation content below 0.1% by weight, which comprises:
- hydrolyzing a C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin with sodium hydroxide in an amide solvent in the absence of water, optionally in combination with ether;
  - removing said amide solvent to obtain a solid residue of rosuvastatin sodium;
  - optionally washing the obtained solid residue of rosuvastatin sodium with an ether and isolating solid rosuvastatin sodium;
  - combining the solid residue of rosuvastatin sodium with water to obtain an aqueous solution of rosuvastatin sodium;
  - adding a source of calcium to the resulted solution to precipitate rosuvastatin calcium;
  - redigesting the obtained rosuvastatin calcium in water,
  - stirring the obtained rosuvastatin calcium with a dispersion making device with high speed rotation unit; and
  - isolating amorphous rosuvastatin calcium salt having a residual sodium cation content below 0.1% by weight.
9. The process according to claims 8, wherein the C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin is methyl rosuvastatin or *tert*-butyl rosuvastatin.
10. The process according to claim 9, wherein the C<sub>1</sub> to C<sub>5</sub> alkyl ester of rosuvastatin is *tert*-butyl rosuvastatin.
11. The process according to claim 8, wherein the amide solvent is N,N-dimethylacetamide, optionally in combination with ether.
12. The process according to claim 8 or claim 11, wherein the ether is diethyl ether.
13. The process according to claim 12, wherein the amide solvent is N,N-dimethylacetamide, and wherein the ratio of N,N-dimethylacetamide and diethyl ether is from 1:1 to 3:1 by volume.
14. The process according to claim 8, wherein the source of calcium is calcium chloride, calcium nitrate, calcium hydroxide or a calcium salt of a C<sub>1</sub> to C<sub>20</sub> alkanoic acid.

15. The process according to claim 14, wherein the calcium salt of a C<sub>1</sub> to C<sub>20</sub> alkanoic acid is calcium palmitate, calcium pivalate, calcium acetate or calcium acetate monohydrate.
16. The process according to claim 1 or 8, wherein the operation speed of the dispersion making device with high speed rotation unit is from 6000 to 30000 rpm.
17. The process according to claim 16, wherein the operation speed of the dispersion making device with high speed rotation unit is from 15000 to 25000 rpm.
18. The process according to claim 17, wherein the operation speed of the dispersion making device with high speed rotation unit is from 16000 to 20000 rpm.
19. A process for preparing a pharmaceutical formulation comprising rosuvastatin calcium, said process comprising preparing rosuvastatin calcium by the process according to any one of claims 1 to 18, and formulating said rosuvastatin calcium into a pharmaceutical formulation comprising rosuvastatin calcium.



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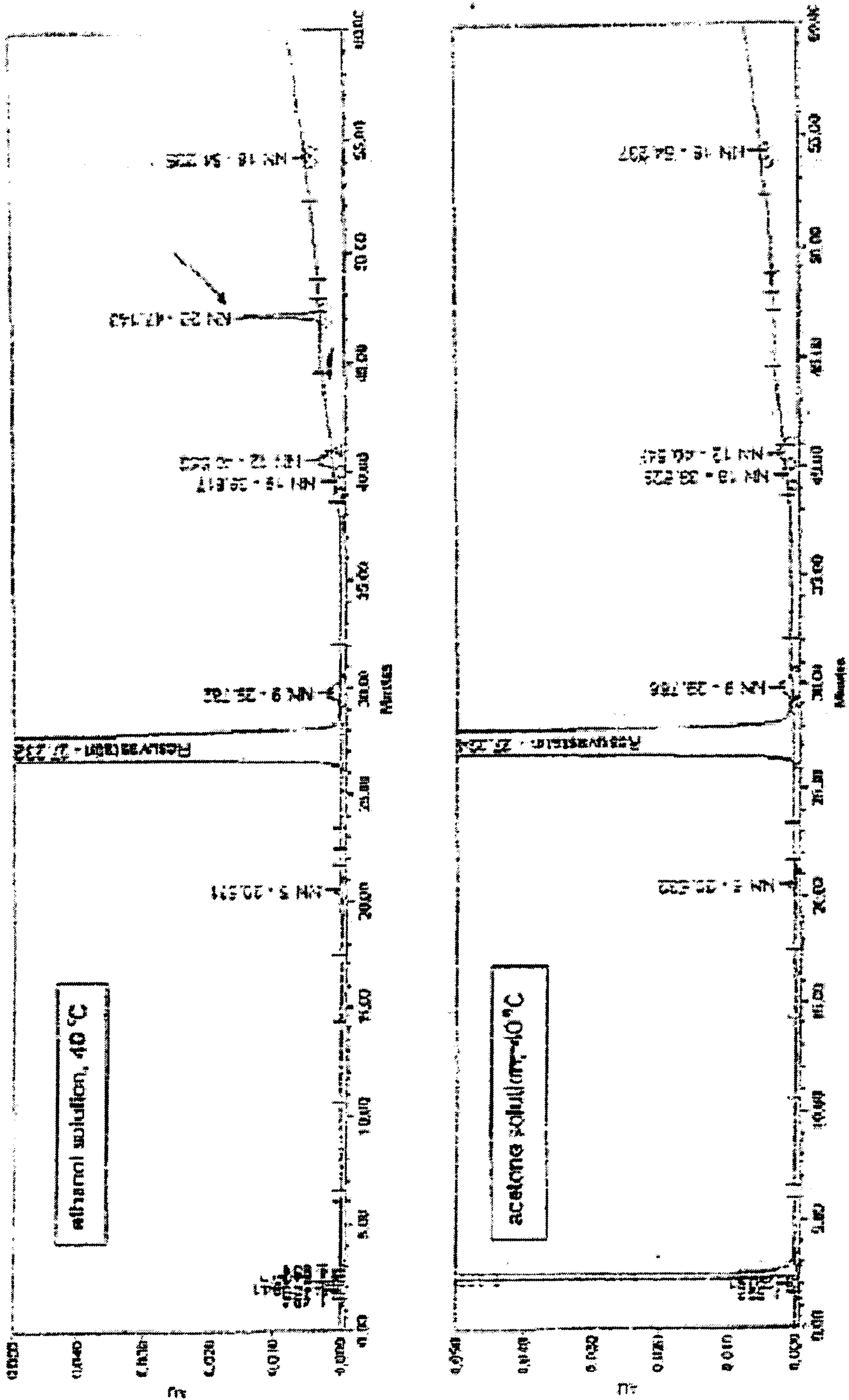


Figure 1

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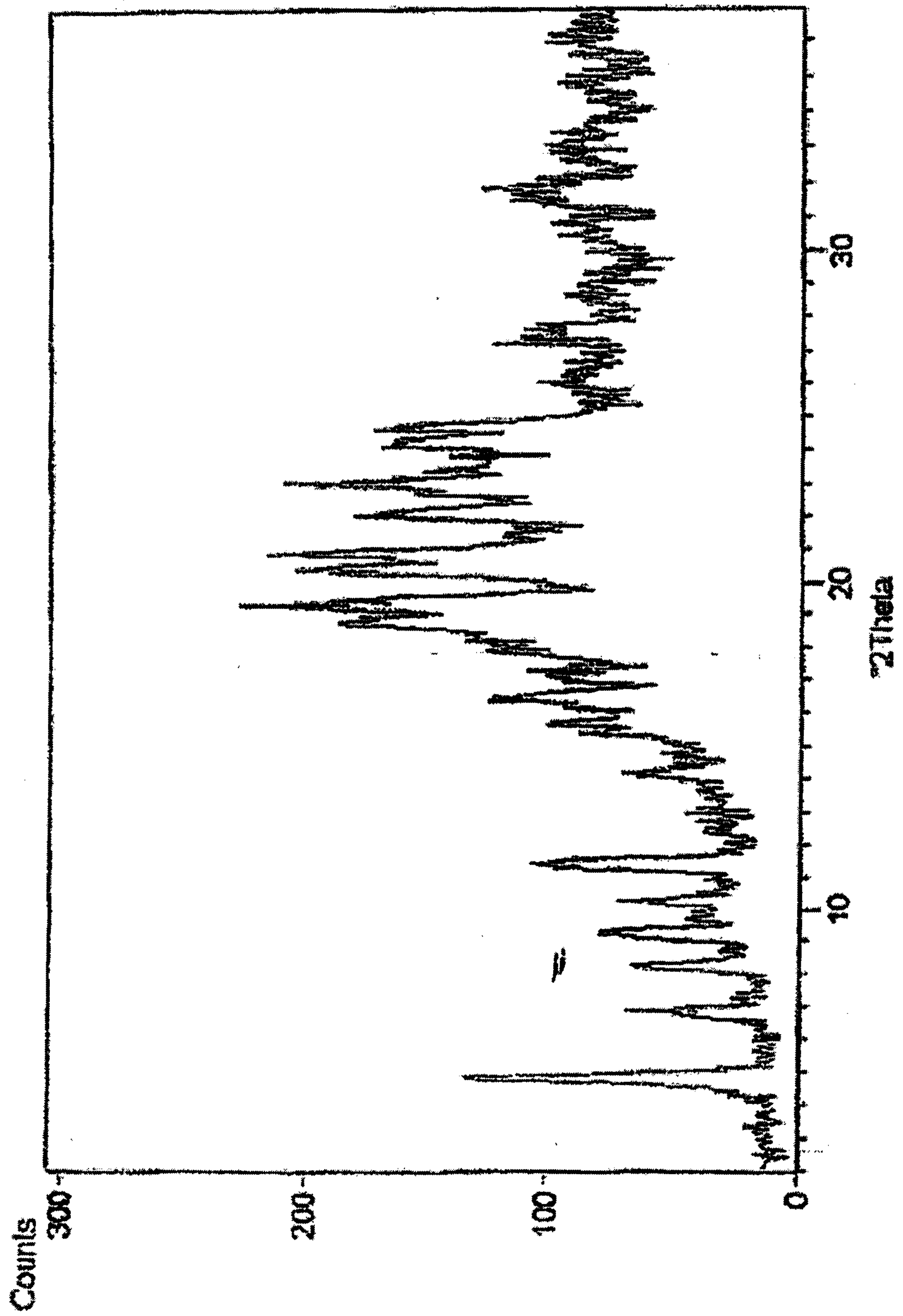
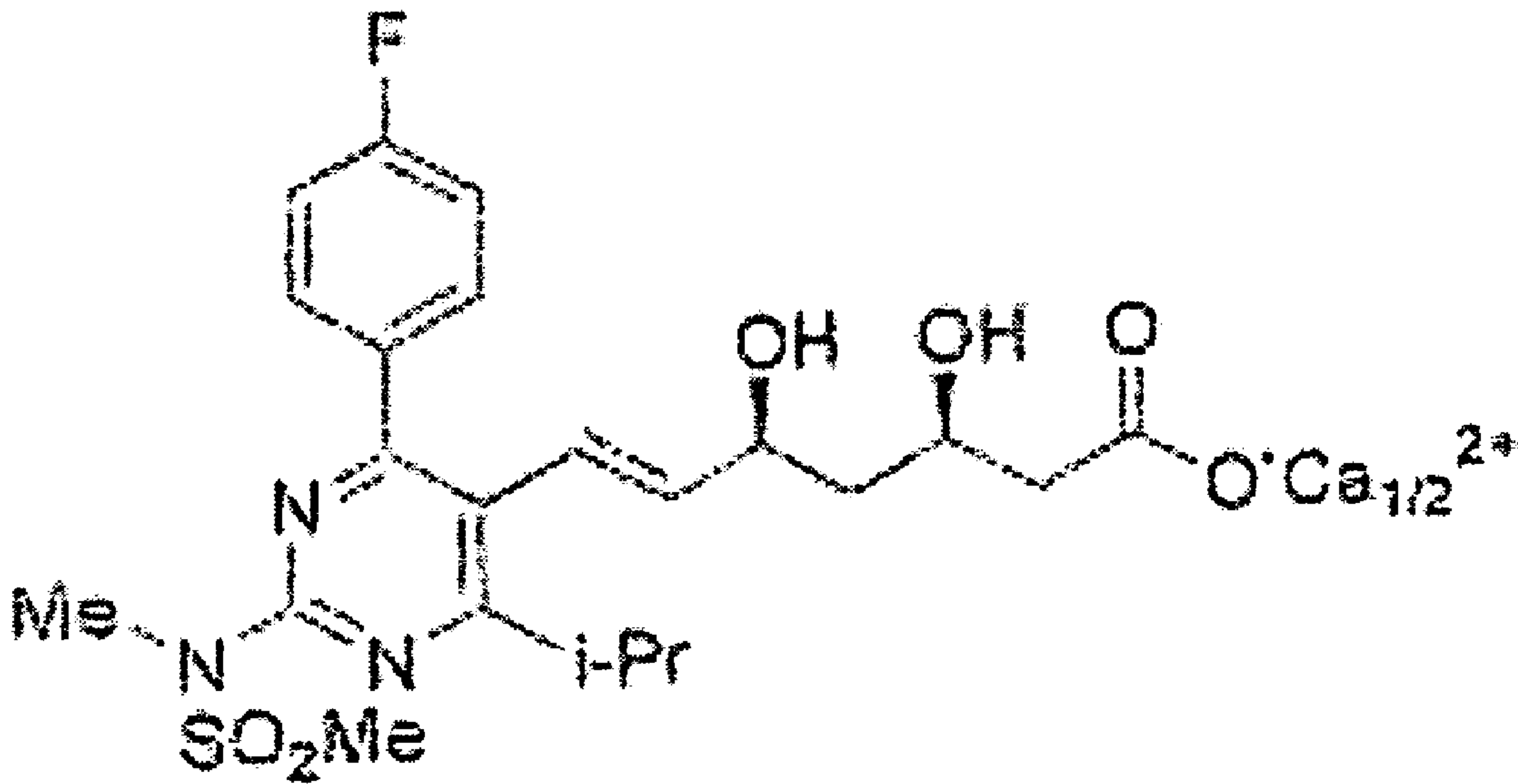


Figure 2.



Formula 1