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(54) **Non-lead sealing glasses** /

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Verres de scellement sans plomb

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DescriptionFIELD OF THE INVENTION

5 The invention relates to tin phosphate glasses used as sealing frits in vitreous fusion seals.

Numerous special sealing glasses have been developed for use in joining glass parts with each other, or with metals, alloys, or ceramics.

10 In making a fusion-type seal, the material must be heated to a temperature where it becomes soft enough to wet the sealing surface and form an adherent, hermetic bond, while maintaining the sealing temperature as low as possible, particularly for sealing electrical and electronic articles.

Low temperature sealing lead glasses, having softening points in the 430°-500°C range and coefficients of thermal expansion in the $70-90 \times 10^{-7}/^{\circ}\text{C}$ range, are described, e.g. in U.S. Patent No. 2,642,633. Lead-zinc borate glasses thermally devitrifying or crystallizing have been investigated for cathode ray tube (CRT) sealing materials.

15 For this and other purposes, the glasses are used as powders or frits, commonly mixed with an organic vehicle, such as amyl acetate, to form a flowable or extrudable paste, this mixture being applied to the sealing surface, optionally with mill additions to modify and/or control thermal expansion of the seal.

For electrical and electronic sealing purposes, even lower sealing temperatures are needed.

To achieve low residual strain in a seal, its coefficient of thermal expansion (CTE) should closely match the CTEs of the parts being sealed.

20 Electronic and electrical sealing applications tend to involve a large number of different materials. Further, these materials tend to have CTEs that vary over a wide range. The required CTEs are generally lower than the values of $95-105 \times 10^{-7}/^{\circ}\text{C}$ encountered in sealing CRT components, e.g. to match alumina; 30-40 to match silicon and borosilicate glasses; and 0-10 to match certain glass-ceramics and fused silicas (always $\times 10^{-7}/^{\circ}\text{C}$).

25 In addition to flow and expansion compatibility, a sealing glass frit desirably has good wettability of the glass parts being sealed, solubility in a common industrial solvent for salvage purposes, and compatibility with organic vehicles, especially amyl acetate.

There has been a continuing desire for a sealing glass frit having all the favorable characteristics of lead-zinc borate glasses, but with even lower sealing temperature and, for health and safety, without lead.

30 The present invention proposes to provide a lead-free sealing glass frit; particularly suited to form fusion seals between electronic and electrical components having CTEs useful for sealing components with lower CTEs; to provide a mid-temperature sealing glass frit, for fusion sealing at 400-430°C; optionally strengthened by reinforcing fibers and fillers therein.

Phosphate glasses which may contain tin and/or zinc are shown by U.S. Patents 2,400,147, 4,940,677, 5,021,355, 5,071,795, 3,407,091, 4,314,031, 5,089,445, 5,089,446.

SUMMARY OF THE INVENTION

35 The invention resides in lead-free, tin phosphate glasses, the compositions of which, as calculated in mole percent on an oxide basis, consist essentially of 25-50% P_2O_5 , 30-70% SnO , 0-15% ZnO , the mole ratio of $\text{SnO}:\text{ZnO}$ being greater than 5:1, and an effective amount up to 25% total of at least one stabilizing oxide in the indicated proportion selected from the group consisting of up to 25% R_2O , wherein R_2O consists of 0-25% Li_2O , 0-25% Na_2O , and 0-25% K_2O , up to 20% B_2O_3 , up to 5% Al_2O_3 , up to 5% SiO_2 , and up to 5% WO_3 ; sealing materials containing such glasses as an active ingredient; and composite articles with at least two component parts joined by a fusion seal, that is, the fused product of a sealing material containing such glasses, as an active ingredient.

DESCRIPTION OF THE INVENTION

40 Our invention is based in large measure on discovery of a family of tin phosphate glasses. It is further based on the discovery that these glasses provide highly effective low- and mid-temperature (350-450°C) sealing glasses, particularly in frit form. While not limited in their fields of application, these glasses are of particular interest as replacements for the high lead content sealing glasses currently used as mid-temperature sealing glasses. A major feature of the present glasses is their freedom from lead.

45 The glasses disclosed in the Aitken et al. application are characterized by a $\text{SnO}:\text{ZnO}$ ratio of 1:1 to 5:1. The present glasses also preferably contain ZnO , but in smaller amounts with respect to the SnO content. Thus, the presence of ZnO is optional, and, when present, the molar ratio of $\text{SnO}:\text{ZnO}$ is greater than 5:1.

The decrease in relative ZnO content tends to soften the glass and result in good flow characteristics. However, it creates a problem of glass exuding during sealing and creating a weak seal. For example, a stannous pyrophosphate composition, that is, a glass consisting of SnO and P_2O_5 , showed some flow at 345°C, but tended to exude glass, to

devitrify and to become weak at higher sealing temperatures. Exudation of glass is integrally associated with devitrification. To counter this trend, ZnO may be added, but the glasses with high SnO:ZnO ratios are still difficult to control.

We have found that the glasses with higher SnO:ZnO ratios can be stabilized to provide good sealing frits. To this end, we include up to 25 mole % total of at least one oxide selected from the group consisting of up to 25 mole % R_2O , wherein R_2O consists of 0-25% Li_2O , 0-25% Na_2O , and 0-25% K_2O , up to 20% mole % B_2O_3 , up to 5 mole % Al_2O_3 , up to 5% SiO_2 , and up to 5% WO_3 , the selected oxide being present in an amount effective to stabilize the glass against devitrification and/or exudation during sealing. In general, we have found that at least 1 mole % of one of the above oxides alone or a combination of two or more of the oxides is necessary to assure substantial stabilization. With such additions, reasonably good flow characteristics can be retained while avoiding the crystallization and loss of strength encountered in their absence. Li_2O is the most preferred alkali metal oxide (R_2O) for stabilization purposes.

A further problem is encountered in that coefficients of thermal expansion (CTEs) of the present glasses tend to be in the range of $120-140 \times 10^{-7}/^{\circ}C$. As indicated earlier, such values are much too high for many electronic and electrical sealing applications. We have found this can be remedied by adding substantial amounts of certain fillers that provide a lower effective CTE in a sealing frit. Among the fillers found particularly effective are cordierite, $Li_2O-Al_2O_3-SiO_2$ glass-ceramics, and a crystalline pyrophosphate having at least one cation component selected from the group consisting of Mg^{+2} and Co^{+2} .

The capability of halides, particularly fluorides, to soften a glass, and thus lower its sealing temperature, is well known. However, where a product must operate under a vacuum, there is concern that such halide-containing glasses may outgas during bakeout. Therefore, it is a feature of our glasses that they may be halide-free.

It is also known that the presence of alkali metal oxides (R_2O) may soften a glass. Such additives may also improve wettability of a sealing surface. However, they tend to increase the coefficient of thermal expansion and/or lower resistance to weathering. Further, in electrical applications, alkali in a glass may adversely affect electrical properties, and may tend to migrate. Accordingly, it is a further feature of our glasses that they may be alkali-free.

In addition to this primary function of stabilizing the present glasses, up to about five mole percent SiO_2 lowers the coefficient of thermal expansion. Up to about twenty mole percent B_2O_3 , but preferably not over ten, lowers the coefficient of thermal expansion of the glass, and up to about five mole percent Al_2O_3 improves durability.

A key requirement in the present glasses is that the glass be in a reduced state, that is, the tin be predominantly in the stannous (Sn^{+2}) state. To this end, tin is added to a glass batch in stannous form, that is, as black tin (SnO). Alternatively, if light tin (SnO_2) is employed, a reducing agent, such as sugar, should be added to insure that a predominant amount of the tin is in a divalent (SnO) state. However, care must be taken not to employ such vigorous reducing conditions that the tin is further reduced to the metal. Where SnO_2 is present in substantial amount, the frit does not flow and wet the sealing surface as desired for a strong seal.

Good glass formation is obtained with P_2O_5 contents ranging from an orthophosphate level, i.e., about 25 mole % P_2O_5 , to a metaphosphate level, i.e., about 50 mole % P_2O_5 .

For a frit to use in sealing components, we prefer that the glass contains 29-35 mole percent P_2O_5 , more preferably at or near the pyrophosphate stoichiometry, that is, about 33% P_2O_5 . Lower P_2O_5 content frits tend to provide erratic and nonreproducible flow behavior in sealing operations. Frits with higher P_2O_5 levels tend to be less durable against chemical attack.

Optionally, we can include at least one component selected from the group consisting of up to 5 mole % MoO_3 , up to 5 mole % WO_3 , and up to 0.1 mole % Ag in the glass batch, the selected component being present in an amount effective to promote the adherence characteristics of the sealing frit. Those maxima correspond to 5.5 wt. % MoO_3 , 8.7 wt. % WO_3 and 0.08 wt. % Ag metal.

The distinctive nature of the present glass sealing frits, as compared to those described in the Aitken et al. application noted earlier, may be seen from comparison of certain relevant properties. All glasses compared were three-component glasses ($SnO-ZnO-P_2O_5$) with 33 mole percent P_2O_5 and differing only in SnO:ZnO ratios.

Glass transition temperature (T_g) is a measure of the temperature at which flow can initiate in a glass as determined by differential scanning calorimetry (DSC). T_g values decrease rapidly from about $450^{\circ}C$ to about $285^{\circ}C$ as the SnO:ZnO ratio is increased up to 5:1 from 0, i.e., a zinc pyrophosphate. As the ratio is further increased (the present glasses), the T_g value stabilizes in the range of $250-275^{\circ}C$.

In like manner, the glass annealing point drops rapidly from about $270^{\circ}C$, at a SnO:ZnO ratio of about 1:1, to about $245^{\circ}C$, and then stabilizes as the SnO level is further increased. However, it is nearly mutually contradictory for a frit to possess both a low-flow-temperature and a low expansion coefficient (CTE). The $SnO-ZnO-P_2O_5$ glasses are no exception. Consequently, the present low ZnO content glasses have CTE values in the range of $120-140 \times 10^{-7}/^{\circ}C$ over the temperature interval of room temperature (R.T.) to $250^{\circ}C$.

The invention is further described with reference to glass compositions and sealing mixtures that illustrate the invention.

The compositions in TABLE IA are reported in terms of mole percent on the oxide basis as calculated from a glass batch. Corresponding compositions in terms of parts by weight are presented in TABLE IB. Since the totals of these compositions approximate 100, the individual values may be taken as weight percentages.

TABLE IA shows a series of basic, three-component, SnO-ZnO-P₂O₅ glass compositions (in mole percent) with increasing SnO:ZnO ratios. Also shown is the temperature (Temp°C) at which a pressed frit glass cylinder or button exhibited excellent flow. It should be noted, however, that such flow was accompanied by glass exudation, and a weak, crumbling structure as a test piece cooled. This was more pronounced at the higher SnO:ZnO ratios.

TABLE IA

	1	2	3	4	5
P ₂ O ₅	33.0	33.0	33.0	33.0	33.0
ZnO	11.2	7.4	5.0	3.2	0.0
SnO	55.8	59.6	62.0	63.8	67.0
SnO/ZnO	5.0	8.0	12.4	20.0	∞
Temp (°C)	430	410	410	390	350

TABLE IB

	1	2	3	4	5
P ₂ O ₅	35.7	35.2	34.9	34.6	34.2
ZnO	7.0	4.5	3.0	1.9	--
SnO	57.3	60.3	62.1	63.5	65.8

All batches were melted in covered silica crucibles. Typical batches were 800-1000g. The melting temperatures were in the range of 800°-1100°C, depending on the glass composition. Melting time was generally held constant at 2 hours. The batch raw materials consisted of SnO (black tin), zinc oxide, and ammonium phosphate. Glasses were formed by pouring the melts onto a steel table, and cooling the melts using a roller to form 0.1" (~0.25 cm) thick ribbon. All compositions were then prepared as frits by dry ball-milling to a mean particle size of 20-50 μm. Samples for thermal expansion or viscosity were obtained by dry-pressing bars from these powders, and then sintering. The bars were fired to a temperature high enough that sufficient densification occurred so that, after cooling, test specimens could be sawed from the sintered bars. Typically, a 1 hour soak time at maximum temperature was used for all specimens.

It is clear from TABLE IA that substantial benefits are associated with higher SnO/ZnO ratio frits since flow can occur at temperatures as low as 350°C. It is notable that these low flow temperatures are obtained without the use of alkali oxides or halides. This makes these frits very attractive for electronic sealing applications where the presence of either alkalis or halides is undesirable because of contamination and possible degradation in electrical stability of the sealed device. In addition, the low sealing temperatures for this family of frits do not require the use of lead or thallium oxide, two components which are toxic and/or hazardous, and typically present in low temperature frits.

However, as noted earlier, frits with high SnO/ZnO ratios have a tendency to crystallize (especially at soak times >0.5hr) in such a fashion that any residual glassy phase is exuded to the surface. This is extremely undesirable for the use of these frits for sealing purposes. In addition, the absence of any glassy phase within the interior of the frit produces a very porous and weak structure.

The glass exudation/crystallization phenomena may be eliminated by composition modification of the glass, principally by the addition of Al₂O₃, B₂O₃, or SiO₂. These additional oxides presumably function to depress the tendency for crystallization by tightening the phosphorus-oxygen chains in the glass, making it more difficult for structural rearrangements to take place leading to crystallization.

The effects of B₂O₃ and Al₂O₃ on crystallization and glass exudation are presented in TABLE IIA for a series of frits with SnO/ZnO ratios in the range 10.0 to ∞, i.e., a tin pyrophosphate. Data are shown for three different SnO/ZnO ratios for "simple" three component frits (i.e., SnO, ZnO, and P₂O₅), and for the same compositions with small amounts of Al₂O₃, B₂O₃, and/or Li₂O or WO₃ added to the glass, all on a mole basis. In each instance, these additional components served to suppress the tendency for glass exudation (exud.).

TABLE IIA

Ex.	6	7	8	9
SnO/ZnO	10.0	10.0	12.4	12.4
P ₂ O ₅	33.0	32.0	33.0	31.8
ZnO	6.1	5.9	5.0	4.8
SnO	60.9	59.0	62.0	59.8
B ₂ O ₃	----	1.9	----	2.4
Al ₂ O ₃	----	0.6	----	1.2
WO ₃	----	0.5	----	----
Li ₂ O	----	----	----	----
<u>Firing Schedule (1h hold)</u>	380°	380°	430°	430
<u>Appearance</u>	exud., crumbly	glossy, no exud.	exud., crumbly	glossy, no exud.
Ex.	10	11	12	13
SnO/ZnO	20.0	20.0	∞	∞
P ₂ O ₅	33.0	32.6	33.0	31.4
ZnO	3.2	2.5	----	----
SnO	63.8	49.7	67.0	63.7
B ₂ O ₃	----	----	----	3.8
Al ₂ O ₃	----	1.0	----	1.0
WO ₃	----	0.5	----	----
Li ₂ O	----	13.8	----	----
<u>Firing Schedule (1h hold)</u>	410°	410°	380°	380°
<u>Appearance</u>	exud., crumbly	glossy, no exud.	exud., crumbly	glossy, no exud.

TABLE IIB records the compositions of TABLE IIA in terms of parts by weight. In like manner to TABLE IB, the totals of the compositions so closely approximate 100 such that, for all practical purposes, the tabulated value of each constituent may be deemed to reflect weight percent.

TABLE IIB

	6	7	8	9	10	11	12	13
P ₂ O ₅	35.0	34.2	34.9	34.1	34.5	38.1	34.2	33.5
ZnO	3.7	3.6	3.0	2.9	1.9	1.7	--	--
SnO	61.3	59.8	62.1	60.8	62.6	55.1	65.8	63.8
B ₂ O ₃	--	1.0	--	1.3	--	--	--	2.0
Al ₂ O ₃	--	0.5	--	0.9	--	0.8	--	0.8
WO ₃	--	0.9	--	--	--	0.9	--	--
Li ₂ O	--	--	--	--	--	3.4	--	--

The SnO-ZnO-P₂O₅ frits described in this invention disclosure are vitreous, and have a low elastic modulus [approximately 6x10⁶ psi (~4.2 x 10⁴ MPa)]. Consequently, these frits have relatively low strength values when tested

on a pressed bar or disk fired to a temperature at which flow occurs. Typical strength values for sintered disks are 3,000-4,000 psi (~21 - 28 MPa), compared to 5,000-5,500 psi (~35 - 38.5 MPa) for a commercial, vitreous, PbO-containing frit. However, as noted previously, the low flow temperatures of these frits permit them to accommodate relatively large amounts of fillers (added as mill additions) and still flow at 400°-500°. The low Young's modulus is actually beneficial since it permits the fillers to be accommodated in the frit without microcracking and resulting strength degradation.

For strength testing, fillers were mixed with the powdered frit by dry-mixing on a set of rollers. The various powders were placed in a plastic container along with some grinding media, and then roller-mixed for 15-30 minutes. Strength specimens were prepared by pressing these mixed powders into disks approximately 1.5" diameter x 0.20" thick (~3.8 x 0.5 cm). After firing, thin disks were prepared by core-drilling and grinding. One face (the tensile surface) was polished, and then, just prior to actual strength testing, was abraded to introduce a uniform flaw distribution.

The fillers that were investigated were -325 mesh Al_2O_3 , (estimated CTE = 70), E-mill zircon, (estimated CTE = 50), and S grade zirconia, (estimated CTE = 150). The mean particle size of all these fillers was in the 5-10 μm range. Of the materials evaluated as fillers, only $\alpha-Al_2O_3$ and zircon were effective in improving strength. Presumably, this was because their expansion coefficient was lower than that of the frit (CTE of frit = 120-140). Cracks tend to deflect towards, and interact with, second phase inclusions having a CTE lower than that of the surrounding matrix. It is thus likely that other fillers with CTE <120, or preferably with CTE <100, would also be effective as reinforcement agents.

Examples of reinforced frits, along with measured strengths, are presented in TABLE III. Generally, 8-10 specimens were tested per set. All specimens were fired at 450°C or lower. Because of the desire to keep potential seal temperatures < 450°C, filler loadings were maintained at 40 wt% or less. Strength values in excess of 5,000 psi (~35 MPa) were obtained for all filler combinations involving Al_2O_3 and/or zircon at filler loadings > 25 wt%. Certain combinations of Al_2O_3 and zircon produced strength values in excess of 8,000 psi (~56 MPa).

TABLE III

Strength Properties of Reinforced Frits		
Frit (mole%) Composition	Batch Recipe (wt %) Reinforced Frit	Strength, Mean \pm S.D. psi
P_2O_5 33.0	80% frit-20% ZrO_2	3680 \pm 180 (~25.8MPa)
ZnO 13.4	70% frit-30% ZrO_2	4300 \pm 400 (~30.1MPa)
SnO 53.6	60% frit-40% ZrO_2	3600 \pm 430 (~25.2MPa)
SnO/ZnO 4.0	80% frit-20% zircon	4640 \pm 360 (~32.5MPa)
	70% frit-30% zircon	6000 \pm 370 (~42 MPa)
	60% frit-40% zircon	5780 \pm 370 (~40.5MPa)
	65% frit-25% zircon - 10% Al_2O_3	8080 \pm 470 (~56.6MPa)
P_2O_5 32.0	80% frit-20% Al_2O_3	5520 \pm 230 (~38.6MPa)
ZnO 9.3	70% frit-30% zircon	7350 \pm 760 (~51.5MPa)
SnO 55.7	65% frit-25% zircon - 10% Al_2O_3	8080 \pm 470 (~56.6MPa)
Al_2O_3 0.6		
B_2O_3 1.9		
SnO/ZnO 6.0		

The higher SnO/ZnO ratio frits, although they have low flow temperatures, also have relatively high thermal expansion coefficients, generally in the range 120-140. This is well above values desirable for sealing to potential substrates such as soda-lime-silica glass (CTE = 90), alumina (CTE = 70), ZnO (CTE = 55), borosilicate glass (CTE = 35), and silicon (CTE = 30). However, the good flow properties of these glasses permit relatively large amounts of mill additions to be added while still keeping potential sealing temperatures within the moderate range of 400°-500°. Potential mill additions include lithium aluminosilicate glass ceramic (LAS), cordierite, and crystalline magnesium pyrophosphates. Listed in TABLE IV are several examples of SnO-ZnO- P_2O_5 frits made with mill additions of expansion modifiers. Also given are sealing temperature and the particular substrate used.

TABLE IV

Frit (mole %) Composition	Batch Recipe (wt%)	Sub- strate	Seal Temp	Results
P ₂ O ₅ 32.0	70% frit-20% LAS	Al ₂ O ₃	430°	Good adherence Good flow
ZnO 9.3				
SnO 55.7				
Al ₂ O ₃ 0.6				
B ₂ O ₃ 1.9				
SnO/ZnO 6.0				
P ₂ O ₅ 32.0	70% frit-30% LAS	ZnO	550°	Good adherence Good flow
ZnO 9.3				
SnO 55.7				
Al ₂ O ₃ 0.6				
B ₂ O ₃ 1.9				
SnO/ZnO 6.0				

TABLE IV Cont.

	Frit (mole %)		Batch	Sub-	Seal	Results
	Composition		Recipe (wt%)	strate	Temp.	
10	P ₂ O ₅	32.0	67% frit-28% cordierite-5% Magnesium pyrophosphate	boro- silicate glass	500°	Moderate adherence Poor flow
	ZnO	9.3				
	SnO	55.7				
15	Al ₂ O ₃	0.6				
	B ₂ O ₃	1.9				
20	SnO/ZnO	6.0				
	P ₂ O ₅	34.0	70% frit-30% LAS	Al ₂ O ₃	400°	Good adherence Moderate flow
25	ZnO	3.0				
	SnO	59.7				
	Al ₂ O ₃	0.8				
30	B ₂ O ₃	2.0				
	WO ₃	0.5				
35	SnO/ZnO	20.0				
	P ₂ O ₅	34.0	67% frit-33% LAS	boro- silicate glass	500°	Good adherence Good flow
40	ZnO	3.0				
	SnO	59.7				
	Al ₂ O ₃	0.8				
45	B ₂ O ₃	2.0				
	WO ₃	0.5				
50	SnO/ZnO	20.0				

TABLE IV Cont.

Frit (mole %) Composition	Batch Recipe (wt%)	Sub- strate	Seal Temp.	Results
P ₂ O ₅ 34.0	67% frit-28% LAS, 5%	boro- silicate	500°	Good adherence Good flow
ZnO 3.0	Magnesium pyrophosphate	glass		
SnO 59.7				
Al ₂ O ₃ 0.8				
B ₂ O ₃ 2.0				
WO ₃ 0.5				
SnO/ZnO 20.0				

Claims

1. A lead-free, tin phosphate glass, the composition of which, as calculated in mole percent on an oxide basis, consists essentially of 25-50% P₂O₅, 30-70% SnO, 0-15% ZnO, the mole ratio of SnO:ZnO being greater than 5:1, and an effective amount up to 25% total of at least one stabilizing oxide in the indicated proportion selected from the group consisting of up to 25% R₂O, wherein R₂O consists of 0-25% Li₂O, 0-25% Na₂O, and 0-25% K₂O, up to 20% B₂O₃, up to 5% Al₂O₃, up to 5% SiO₂, and up to 5% WO₃.
2. A phosphate glass in accordance with claim 1 wherein the selected stabilizing oxide is B₂O₃ and/or Al₂O₃.
3. A phosphate glass in accordance with claim 1 wherein the P₂O₅ content is 29-35 mole percent.
4. A phosphate glass in accordance with claim 3 wherein P₂O₅ content is about 33 mole percent.
5. A phosphate glass in accordance with claim 1 which also contains an effective amount of at least one seal adherence promoter in the indicated proportion selected from the group consisting of up to 5 mole % WO₃, up to 5 mole % MoO₃, and up to 0.10 mole % Ag metal.
6. Glass according to any one of claims 1-5 additionally containing a mill addition having a low coefficient of thermal expansion to reduce the effective coefficient of thermal expansion of the frit in a fusion seal.
7. Glass according to claim 6, wherein the mill addition is selected from the group consisting of lithium aluminosilicate glass-ceramics, cordierite, and a pyrophosphate having at least one cation selected from the group consisting of Mg⁺² and Co⁺², or particles of a magnesium pyrophosphate crystalline material.
8. Glass according to any one of claims 1-7, also containing a strength reinforcing filler having a coefficient of expansion less than $120 \times 10^{-7}/^{\circ}\text{C}$.
9. Glass according to claim 8 wherein the strength reinforcing filler is selected from the group composed of alumina and zircon.

10. Glass according to any one of the preceding claims, used as active ingredient in a sealing material, or in a composite article composed of at least two component parts joined by a fusion seal provided by said fused sealing material.

Patentansprüche

1. Bleifreies Zinnphosphatglas mit einer Zusammensetzung, berechnet in Mol-% auf Oxidbasis, die im wesentlichen aus 25-50 % P_2O_5 , 30-70 % SnO , 0-15 % ZnO , wobei das Mol-Verhältnis $SnO:ZnO$ größer als 5:1 ist, und einer wirksamen Menge von bis zu insgesamt 25 % wenigstens eines stabilisierenden Oxids im angegebenen Anteil, ausgewählt aus der Gruppe, bestehend aus bis zu 25 % R_2O , wobei R_2O aus 0-25 % Li_2O , 0-25 % Na_2O und 0-25 % K_2O besteht, bis zu 20 % B_2O_3 , bis zu 5 % Al_2O_3 , bis zu 5 % SiO_2 und bis zu 5 % WO_3 , besteht.
2. Phosphatglas nach Anspruch 1, wobei das ausgewählte stabilisierende Oxid B_2O_3 und/oder Al_2O_3 ist.
3. Phosphatglas nach Anspruch 1, wobei der P_2O_5 -Gehalt 29 - 35 Mol-% beträgt.
4. Phosphatglas nach Anspruch 3, wobei der P_2O_5 -Gehalt etwa 33 Mol-% beträgt.
5. Phosphatglas nach Anspruch 1, das weiterhin eine wirksame Menge wenigstens eines Dichtungsdhäsionspromotors im angegebenen Anteil enthält, ausgewählt aus der Gruppe, bestehend aus bis zu 5 Mol-% WO_3 , bis zu 5 Mol-% MoO_3 und bis zu 0,10 Mol-% Ag-Metall.
6. Glas nach einem der Ansprüche 1 - 5, das zusätzlich einen Vermahlungsmittelzusatz mit einem niedrigen Wärmeausdehnungskoeffizienten enthält, um den wirksamen Wärmeausdehnungskoeffizienten der Fritte in einer Schmelzdichtung herabzusetzen.
7. Glas nach Anspruch 6, wobei der Vermahlungsmittelzusatz aus der Gruppe, bestehend aus Lithiumaluminosilikat-Glaskeramik, Cordierit und einem Pyrophosphat mit wenigstens einem Kation, ausgewählt aus der Gruppe, bestehend aus Mg^{+2} und Co^{+2} , oder Teilchen aus einem kristallinen Magnesiumpyrophosphatmaterial ausgewählt ist.
8. Glas nach einem der Ansprüche 1 - 7, weiterhin enthaltend einen Festigkeitsverstärkungsfüllstoff mit einem Ausdehnungskoeffizienten von unter $120 \times 10^{-7}/^{\circ}C$.
9. Glas nach Anspruch 8, wobei der Festigkeitsverstärkungsfüllstoff aus der Gruppe, bestehend aus Aluminiumoxid und Zirkon, ausgewählt ist.
10. Glas nach einem der vorhergehenden Ansprüche, das als Wirkstoff in einem Dichtungsmaterial oder in einem Verbundgegenstand verwendet wird, Zusammengesetzt aus wenigstens zwei Bestandteilen, die durch eine durch das geschmolzene Dichtungsmaterial bereitgestellte Schmelzdichtung verbunden sind.

Revendications

1. Verre de phosphate d'étain sans plomb, dont la composition consiste essentiellement, calculée en pourcents en moles sur la base des oxydes, en 25 à 50 % de P_2O_5 , 30 à 70 % de SnO , 0 à 15 % de ZnO , le rapport molaire de $SnO:ZnO$ étant supérieur à 5:1, et une quantité efficace, représentant jusqu'à 25 % du total, d'au moins un oxyde stabilisant, dans les proportions indiquées, choisi dans l'ensemble constitué par jusqu'à 25 % de R_2O , R_2O consistant en 0 à 25 % de Li_2O , 0 à 25 % de Na_2O et 0 à 25 % de K_2O , jusqu'à 20 % de B_2O_3 , jusqu'à 5 % de Al_2O_3 , jusqu'à 5 % de SiO_2 et jusqu'à 5 % de WO_3 .
2. Verre phosphaté selon la revendication 1, dans lequel l'oxyde stabilisant choisi est B_2O_3 et/ou Al_2O_3 .
3. Verre phosphaté selon la revendication 1, dans lequel la teneur en P_2O_5 est de 29 à 35 % en moles.
4. Verre phosphaté selon la revendication 3, dans lequel la teneur en P_2O_5 est d'environ 33 % en moles.
5. Verre phosphaté selon la revendication 1, contenant aussi une quantité efficace d'au moins un promoteur d'adhérence de scellement, dans les proportions indiquées, choisi dans l'ensemble constitué par jusqu'à 5 % en moles de WO_3 , jusqu'à 5 % en moles de MoO_3 et jusqu'à 0,10 % en moles de métal Ag.

6. Verre selon l'une quelconque des revendications 1 à 5, contenant, en plus, un additif broyé ayant un faible coefficient de dilatation thermique, de manière à réduire le coefficient effectif de dilatation thermique de la fritte dans un scellement par fusion.
- 5 7. Verre selon la revendication 6, dans lequel l'additif broyé est choisi dans l'ensemble constitué par les vitrocéramiques d'aluminosilicate de lithium, la cordiérite et un pyrophosphate ayant au moins un cation choisi dans l'ensemble constitué par Mg^{+2} et Co^{+2} , ou des particules d'un matériau cristallin de pyrophosphate de magnésium.
- 10 8. Verre selon l'une quelconque des revendications 1 à 7, contenant aussi une charge renforçant la résistance mécanique, ayant un coefficient de dilatation inférieur à $120 \times 10^{-7}/^{\circ}C$.
9. Verre selon la revendication 8, la charge renforçant la résistance mécanique étant choisie dans l'ensemble constitué par l'alumine et l'oxyde de zirconium.
- 15 10. Verre selon l'une quelconque des revendications précédentes, utilisé comme ingrédient actif dans un matériau de scellement ou dans un article composite formé d'au moins deux parties composantes réunies par un scellement par fusion fourni par ledit matériau de scellement fondu.

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Title NON-LEAD SEALING GLASSES.

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