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# DESCRIPTION

## FIELD OF THE INVENTION

**[0001]** The present invention relates to systems and methods for detecting target analytes in a sample. In particular, the present invention provides a local plasmon resonance-based analyte detection system capable of detecting a minute quantity of a target analyte in a sample.

## BACKGROUND OF THE INVENTION

**[0002]** Current immunoassays and biomolecule binding assays typically require multiple steps and sophisticated equipment to perform the assays. The lack of sensitivity and the complexity involved in performing such heterogeneous assays arises from the specific need to separate labeled from unlabeled specific binding partners.

**[0003]** Attempts to develop assays based on the local surface plasmon resonance (LSPR) properties of noble metal nanoparticles have been made (Tokel et al., Chem Rev., Vol. 114: 5728-5752, 2014). LSPR is the collective oscillation of electrons in nanometer-sized structures induced by incident light. Metallic nanoparticles have a strong electromagnetic response to refractive index changes in their immediate vicinity and thus shifts in the resonance frequency of the nanoparticles can be measured as an indicator of molecules binding to the nanoparticle surface. Although metallic nanoparticles, particularly gold nanoparticles, have been employed in diagnostic assays to detect binding events, such assays generally suffer from low sensitivity and cannot be used to quantitatively monitor the kinetics of sequential binding events.

**[0004]** Thus, improved assay methods employing a homogenous format while providing increased sensitivity are needed. Assays utilizing standard laboratory techniques, such as spectroscopy, would also be desirable.

## SUMMARY OF THE INVENTION

**[0005]** The present invention is based, in part, on the discovery that composite metallic nanostructures can enhance the optical signals induced by binding of a molecule to a metallic nanolayer surface. The amplification observed greatly increases the sensitivity of detecting specific biomolecular binding events such that sub-picogram quantities of the biomolecule can be detected. Accordingly, the present invention provides analyte detection devices and methods of using such devices to detect minute quantities of a target analyte in a sample.

**[0006]** In one embodiment, the analyte detection devices comprise a plurality of detection conjugates, a surface containing a metallic nanolayer, and a plurality of capture molecules,

wherein the capture molecules are immobilized on the metallic nanolayer and are capable of specifically binding to the target analyte. In embodiments in which the analyte detection devices are configured in a sandwich assay format, the detection conjugates comprise composite metallic nanostructures coupled to binding partners that are capable of specifically binding to the target analyte. In embodiments in which the analyte detection devices are configured in a direct competitive assay format, the detection conjugates comprise composite metallic nanostructures coupled to target analytes.

**[0007]** The composite metallic nanostructures in the detection conjugates generally comprise at least two noble metals, transition metals, alkali metals, lanthanides, or combinations thereof. In some embodiments, the composite metallic nanostructures comprise at least two metals selected from gold, silver, copper, platinum, palladium, cadmium, iron, nickel, and zinc. In certain embodiments, each of the composite metallic nanostructures comprises a core of a first metal and a coating of a second metal. In some embodiments, the core may be silver or copper with a gold coating. In other embodiments, the core of a first metal may be dissolved following coating such that a hollow structure comprised of the second coating metal results.

**[0008]** The metallic nanolayer deposited on the surface can be a metallic film or comprised of a plurality of metallic nanostructures immobilized on the surface. The metallic nanolayer can also be comprised of a noble or transition metal. In some embodiments, the metallic nanolayer comprises gold, silver, copper, platinum, palladium, cadmium, zinc or a composite thereof. In one embodiment, the metallic nanolayer comprises gold. In another embodiment, the metallic nanolayer comprises silver. In still another embodiment, the metallic nanolayer comprises a silver nanolayer overlaid with a gold nanolayer.

**[0009]** The present invention also provides methods of detecting a target analyte in a sample using the analyte detection devices described herein. In one embodiment, the methods comprise mixing the sample with a plurality of detection conjugates, contacting the mixture with a surface containing a metallic nanolayer on which a plurality of capture molecules are immobilized, exposing the surface to a light source at a wavelength range within the ultraviolet-visible-infrared spectrum; and measuring an optical signal from the surface, wherein a change in the optical signal indicates the presence of the target analyte in the sample. In certain embodiments, the methods of the present invention are capable of detecting femtogram to nanogram quantities of a target analyte in sample.

**[0010]** The present invention includes an assay complex comprising a detection conjugate comprising a composite metallic nanostructure coupled to a binding partner; a target analyte; and a metallic nanolayer-coated bead on which a capture molecule is immobilized, wherein the binding partner in the detection conjugate is bound to a first epitope on the target analyte and the capture molecule is bound to a second epitope on the target analyte, thereby forming a complex comprising the detection conjugate, target analyte, and the capture molecule. In some embodiments, the composite metallic nanostructure is a gold-coated silver nanostructure or a gold-coated copper nanostructure and the metallic nanolayer coating on the bead comprises gold.

**[0011]** The disclosure further describes a method for preparing composite metallic nanostructures for use in the detection devices and methods described herein. In one embodiment, the methods comprise preparing a first solution comprising a mixture of a polymer and chloroauric acid, preparing a second solution comprising silver or copper nanostructures, and incubating the first solution with the second solution for a period of time, wherein the resulting mixture comprises gold-coated silver nanostructures or gold-coated copper nanostructures. In certain embodiments, a reducing agent, such as ascorbic acid, is added to the reaction mixture to increase the quantity of nanostructures produced. In one embodiment, the polymer in the first solution is polyvinylpyrrolidone. In another embodiment, the polymer in the first solution is polyvinyl alcohol.

## BRIEF DESCRIPTION OF THE DRAWINGS

**[0012]**

**Figure 1.** Plot of shift in peak wavelength versus acquisition time for a bovine serum albumin (BSA)-coupled gold nanolayer sensor (channel 1) and human IgG-coupled gold nanolayer sensors (channels 2-4). Arrows indicate the injection sequence and concentration of unlabeled protein A, 1 mM HCl, or protein A labeled with colloidal gold (CGC).

**Figure 2.** Plot of shift in peak wavelength versus acquisition time for an anti-CRP C7 antibody-coupled gold nanolayer sensor. Arrows indicate injection sequence of 0 to 100 ng/ml concentrations of CRP in different channels (CRP loading), 1 µg/ml unlabeled anti-CRP C6 antibody, or 3 µg/ml anti-CRP C6 antibody labeled with colloidal gold (C6-CGC). No further C6-CGC binding was observed when sensor surface was occupied with unlabeled anti-CRP C6 antibody.

**Figure 3.** Plot of shift in peak wavelength versus acquisition time for an anti-CRP C7 antibody-coupled gold nanolayer sensor. Arrows indicate the injection sequence of 0 to 100 ng/ml concentrations of CRP in different channels (CRP loading), 1 µg/ml anti-CRP C6 antibody labeled with colloidal gold (C6-CGC), 3 µg/ml C6-CGC, or 1 mM HCl (Acid).

**Figure 4A.** Reflectance spectra of anti-CRP C7 antibody-coupled gold nanolayer sensors loaded with 10 ng/ml CRP at the various C6-CGC concentrations in Figure 3.

**Figure 4B.** Plot of shift in peak wavelength versus acquisition time for an anti-CRP C7 antibody-coupled gold nanolayer sensor incubated with one of three concentrations of CRP following introduction of 3 µg/ml anti-CRP C6 antibody labeled with colloidal gold (C6-CGC). The table at the right depicts peak analysis 700 seconds after introduction of C6-CGC.

**Figure 5.** Plot of shift in peak wavelength versus acquisition time for an anti-CRP C7 antibody-coupled gold nanolayer sensor. Arrows indicate the injection sequence of 0 to 100 ng/ml concentrations of CRP in different channels (CRP loading with incubation time minimized), 3 µg/ml anti-CRP C6 antibody labeled with colloidal gold (C6-CGC), or 1 mM HCl (Acid).

**Figure 6.** Plot of shift in peak wavelength versus acquisition time for traces in Figure 5 following immediate introduction of 3 µg/ml anti-CRP C6 antibody labeled with colloidal gold (C6-CGC). The table at the right depicts peak analysis 700 seconds after introduction of C6-CGC as compared to the peak shifts obtained with CRP incubation (values shown in Figure 4B).

**Figure 7.** Plot of shift in peak wavelength versus acquisition time for an anti-CRP C7 antibody-coupled gold nanolayer sensor incubated with one of three concentrations of CRP and C6 anti-CRP antibody conjugated to gold-coated silver nanostructures. The control was a gold nanolayer sensor with immobilized bovine serum albumin (BSA) in place of C7 antibody.

## DETAILED DESCRIPTION OF THE INVENTION

**[0013]** The present invention is based, in part, on the discovery that significant amplification in LSPR-based assays can be achieved with composite metallic nanostructure-labeled binding partners. Thus, the present invention provides analyte detection devices comprising an LSPR-surface, *e.g.*, a surface containing a metallic nanolayer, a plurality of capture molecules immobilized to the metallic nanolayer, and a plurality of detection conjugates comprising composite metallic nanostructures coupled to biomolecules.

**[0014]** The analyte detection devices can be configured in a sandwich assay format or a direct competitive assay format. For instance, in one embodiment, an analyte detection device in a sandwich assay format comprises (i) a plurality of detection conjugates, wherein the conjugates comprise composite metallic nanostructures coupled to binding partners that are capable of specifically binding to a target analyte, (ii) a surface containing a metallic nanolayer, and (iii) a plurality of capture molecules, wherein the capture molecules are immobilized on the metallic nanolayer and are capable of specifically binding to the target analyte. In another embodiment, an analyte detection device in a direct competitive assay format comprises (i) a plurality of detection conjugates, wherein the conjugates comprise composite metallic nanostructures coupled to target analytes, (ii) a surface containing a metallic nanolayer, and (iii) a plurality of capture molecules, wherein the capture molecules are immobilized on the metallic nanolayer and are capable of specifically binding to the target analytes.

**[0015]** The analyte detection devices of the invention comprise a surface containing a metallic nanolayer. The surface can be any suitable size and shape, such as a chip, a well, a cuvette, or a bead. In some embodiments, the surface is a rectangular chip. In other embodiments, the surface is a disc. In certain embodiments, the surface is the bottom, the cover, and/or interior walls of a cuvette (*e.g.*, cylindrical or rectangular cuvette). In still other embodiments, the surface is an array of non-metallic particles. The surface can be manufactured from various materials including, but not limited to, glass, quartz, silicon, silica, polystyrene, graphite, fabric (*e.g.* polyethylene fabrics), mesh, or a membrane (*e.g.* latex, polyvinyl, nylon, or polyester).

membranes).

**[0016]** A metallic nanolayer is preferably deposited on the surface. The metallic nanolayer may, in some embodiments, cover the entire surface area of the particular surface. In other embodiments, the metallic nanolayer may be deposited only on a portion of the surface. For example, the surface may contain a plurality of depressions or wells and the metallic nanolayer is deposited within the depressions or wells. In other embodiments, the metallic nanolayer may be applied to the surface as a plurality of spaced deposits across the surface. The optical properties of the metallic nanolayer can be adjusted by varying the thickness of the nanolayer and/or the nature of nanostructures. In one embodiment, the nanolayer is comprised of metallic nanoislands. In another embodiment, the nanolayer is comprised of nanorods. Suitable thicknesses of the metallic nanolayer for use in the devices and methods of the invention include from about 0.5 nm to about 100 nm, about 5 nm to about 30 nm, or about 3 nm to about 10 nm. Exemplary surfaces with a metallic nanolayer coating that can be used in the devices and methods of the invention include the surfaces described in U.S. Patent Publication No. 2006/0240573.

**[0017]** In certain embodiments, the metallic nanolayer is a metallic film. Methods of depositing metallic films on a substrate surface are known to those of skill in the art and include, but are not limited to, atomic layer deposition, pulsed laser deposition, drop casting, vapor deposition, and adsorption. See, e.g., Atanasov et al., *Journal of Physics: Conference Series* 514 (2014); Walters and Parkin, *Journal of Materials Chemistry*, 19: 574-590, 2009; and Gupta et al., *J. Appl. Phys.* 92, 5264-5271, 2002. The metallic film may comprise other components, e.g. the metallic film may be a polymer film, a Langmuir-Blodgett film or an oxide film. In some embodiments, the metallic film comprises two layers, wherein each layer comprises a different metal. By way of example, the metallic film may comprise a silver layer overlaid with a gold layer.

**[0018]** In other embodiments, the metallic nanolayer comprises a plurality of metallic nanostructures immobilized to the surface. Metallic nanostructures can be immobilized to the surface by treating the surface material with a reagent to add functional chemical groups, such as cyanide, amine, thiols, carboxyl, aldehyde or maleimide, and reacting the metallic nanostructures with the treated surface. Metallic nanostructures are known to bind to such functional chemical groups with high affinity. In some embodiments, the metallic nanostructures comprising the metallic nanolayer are spherical nanoparticles. Such nanoparticles have diameters that are less than about 300 nm, less than about 200 nm, or less than about 150 nm. In some embodiments, the spherical nanoparticles have a diameter from about 5 nm to about 200 nm, from about 10 nm to about 100 nm, or from about 20 nm to about 60 nm. In certain embodiments, the size of the metallic nanostructures used to create the metallic nanolayer are similar to the size of the composite nanostructures used in the detection conjugates. In such embodiments, matching the size of the two sets of nanostructures can provide an optimal wavelength shift in a reflectance, emission or scattering spectrum.

**[0019]** The metallic nanolayer (metallic film or plurality of metallic nanostructures) may be

composed of a noble metal or composite thereof. In other embodiments, the metallic nanolayer (metallic film or plurality of metallic nanostructures) may be composed of a transition metal or composite thereof. In certain embodiments, the metallic nanolayer comprises a metal selected from gold, silver, copper, platinum, palladium, ruthenium, rhodium, osmium, iridium, titanium, chromium, cadmium, zinc, iron, cobalt, nickel, and composites thereof. In one particular embodiment, the metallic nanolayer (e.g. metallic film or plurality of metallic nanostructures) comprises gold. In another particular embodiment, the metallic nanolayer (e.g. metallic film or plurality of metallic nanostructures) comprises silver. In certain embodiments, the metallic nanolayer (e.g. metallic film or plurality of metallic nanostructures) comprises a composite of gold and silver or gold and copper. Use of alkali metals (e.g., lithium, sodium, potassium, rubidium, cesium, and francium) or lanthanides (e.g., lanthanum, cerium, praseodymium, neodymium, promethium, samarium, europium, gadolinium, terbium, dysprosium, holmium, erbium, thulium, ytterbium, and lutetium) may improve intensity of the LSPR peaks. Accordingly, in some embodiments, the metallic nanolayer (metallic film or plurality of metallic nanostructures) may be composed of one or more alkali metals or lanthanides. In other embodiments, the metallic nanolayer (metallic film or plurality of metallic nanostructures) may be composed of a combination of a noble metal and an alkali metal or lanthanide.

**[0020]** The analyte detection devices of the invention further comprise a plurality of capture molecules immobilized to the metallic nanolayer deposited on a surface. The capture molecules are capable of specifically binding to a target analyte. As used herein, "specific binding" refers to binding to a target molecule with high affinity, e.g., an affinity of at least  $10^{-6}$  M. In some embodiments, the capture molecules are haptens and other small molecules, drugs, hormones, biological macromolecules including, but not limited to, antibodies or fragments thereof (e.g., Fv, Fab, (Fab)<sub>2</sub>, single chain, CDR etc.), antigens, receptors, ligands, polynucleotides, aptamers, polypeptides, polysaccharides, lipopolysaccharides, glycopeptides, lipoproteins, or nucleoproteins. In certain embodiments, the plurality of capture molecules are antibodies. In other embodiments, the plurality of capture molecules are antigens.

**[0021]** Methods of immobilizing molecules to metallic nanolayers or nanostructures are known to those of skill in the art. Such methods include conjugation chemistries, such as those involving 1-Ethyl-3-[3-dimethylaminopropyl]carbodiimide hydrochloride (EDC), sulfo-NHS coupling, hydrophobic binding or thioether chemistry. In some embodiments, the molecule can be coupled to the metallic nanolayer or nanostructure indirectly through a larger carrier molecule or protein. Such indirect coupling is particularly useful when the molecule is small, such as a hormone, a drug, and other small molecules less than 10 kD. Preferably, the carrier protein is not capable of specific interaction with the target analyte.

**[0022]** The analyte detection devices of the invention may also comprise a plurality of detection conjugates. Detection conjugates comprise metallic nanostructures coupled to binding partners capable of specifically binding to a target analyte or the capture molecules depending on the assay configuration. For example, in embodiments in which the device is configured in a sandwich assay format, the detection conjugates comprise metallic



nanostructures coupled or conjugated to binding partners that are capable of specifically binding a target analyte. In other embodiments in which the device is configured in a direct competitive assay format, the detection conjugates comprise metallic nanostructures coupled or conjugated to target analytes.

**[0023]** The binding partners can be the same types of molecules as the capture molecules, including, but not limited to haptens and other small molecules, drugs, hormones, biological macromolecules such as antibodies or fragments thereof (e.g., Fv, Fab, (Fab)<sub>2</sub>, single chain, CDR etc.), antigens, receptors, ligands, polynucleotides, aptamers, polypeptides, polysaccharides, lipopolysaccharides, glycopeptides, lipoproteins, or nucleoproteins. In some embodiments, the binding partners are the same type of molecule as the capture molecules, but preferably bind to the target analyte at a location distinct from the binding site of the capture molecules. By way of example, the binding partners and the capture molecules can both be antibodies that recognize a target analyte, but the epitope to which the binding partners bind the target analyte is separate from and ideally non-overlapping with the epitope to which the capture molecules bind the target analyte. Thus, in certain embodiments, the binding partners are antibodies that recognize a first epitope of a target analyte and the capture molecules are different antibodies that recognize a second epitope of a target analyte.

**[0024]** The metallic nanostructures in the detection conjugates can be composed of a noble metal or composite thereof. In some embodiments, the metallic nanostructures in the detection conjugates may be composed of a transition metal or composite thereof. In some embodiments, the metallic nanostructures in the detection conjugates may comprise an alkali metal or lanthanide in combination with a noble or transition metal. In certain embodiments, metallic nanostructures in the detection conjugates comprise a metal selected from gold, silver, copper, platinum, palladium, ruthenium, rhodium, osmium, iridium, titanium, chromium, cadmium, zinc, iron, cobalt, nickel, and composites thereof. In one embodiment, the metallic nanostructures are gold nanostructures. In another embodiment, the metallic nanostructures are silver nanostructures.

**[0025]** In preferred embodiments, the metallic nanostructures in the detection conjugates are composite metallic nanostructures. "Composite metallic nanostructures" refers to nanostructures that comprise at least two noble metals, transition metals, alkali metals, or lanthanides. The two or more metals may be mixed together, as in an alloy, or the two or more metals may be present in separate portions of the nanostructure. For example, one metal may form the core of the nanostructure, whereas the second metal forms an outer shell or coating of the nanostructure. In some embodiments, the composite metallic nanostructures comprise at least two metals selected from gold, silver, copper, platinum, palladium, ruthenium, rhodium, osmium, iridium, titanium, chromium, cadmium, zinc, iron, cobalt, and nickel. In other embodiments, the composite metallic nanostructures comprise at least two metals selected from gold, silver, copper, platinum, palladium, cadmium, iron, nickel, and zinc. In one particular embodiment, the composite metallic nanostructures comprise gold and silver. In another embodiment, the composite metallic nanostructures comprise gold and copper. In yet another embodiment, the composite metallic nanostructures comprise silver and copper.

**[0026]** In some embodiments, each of the composite metallic nanostructures is an alloy of a first metal and a second metal. In certain embodiments, each of the composite metallic nanostructures comprises a core of a first metal and a coating of a second metal. In one embodiment, the core is silver and the coating is gold. In another embodiment, the core is copper and the coating is gold. In another embodiment, the core is silver and the coating is copper. In some embodiments, each of the composite metallic nanostructures comprises a dielectric core (e.g. silicon dioxide, gold sulfide, titanium dioxide, silica, and polystyrene), a first coating of a first metal, and a second coating of a second metal. In one particular embodiment, the core is silica, the first coating (i.e. inner coating) is a silver coating, and the second coating is a gold coating (i.e. outer coating). In another embodiment, the core is silica, the first coating (i.e. inner coating) is a copper coating, and the second coating is a gold coating (i.e. outer coating).

**[0027]** In some embodiments, the core comprising a first metal is dissolved following the coating process with a second metal to create a hollow structure comprised of the second metal. For instance, coating of a silver core with gold nanoparticles generates a gold shell around the silver core and the silver core is subsequently dissolved or degraded resulting in the formation of a hollow nanogold shell structure.

**[0028]** The metallic nanostructures include spherical nanoparticles as well nanoplates and nanoshells. Nanoplates have lateral dimensions (e.g. edge lengths) that are greater than their thickness. Nanoplates include nanodisks, nanopolygons, nanohexagons, nanocubes, nanorings, nanostars, and nanoprisms. In some embodiments, the metallic nanostructures, including the composite nanostructures, have a geometry selected from spherical nanoparticles, pyramidal nanoparticles, hexagonal nanoparticles, nanotubes, nanoshells, nanorods, nanodots, nanoislands, nanowires, nanodisks, nanocubes, or combinations thereof. Other shapes are also possible, including irregular shapes. In certain embodiments, the size and shape of the metallic nanostructures are not uniform - i.e. the metallic nanostructures are a heterogeneous mixture of different shapes and sizes of nanostructures.

**[0029]** For spherical nanoparticles, suitable diameter ranges include from about 5 nm to about 200 nm, from about 10 nm to about 100 nm, and from about 20 nm to about 60 nm. For nanoplates, edge lengths may be from about 10 nm to about 800 nm, from about 20 nm to about 500 nm, from about 50 nm to about 200 nm, from about 30 nm to about 100 nm, or from about 10 nm to about 300 nm. The thickness of the nanoplates can range from about 1 to about 100 nm, from about 5 nm to about 80 nm, from about 10 nm to about 50 nm, or from about 5 nm to about 20 nm.

**[0030]** In some embodiments, the nanoplates have an aspect ratio greater than 2. The aspect ratio is the ratio of the edge length to the thickness. Preferably, the nanoplates have an aspect ratio from about 2 to about 25, from about 3 to about 20, from about 5 to about 10, from about 2 to about 15, or from about 10 to about 30.

**[0031]** The binding partners or target analytes can be coupled or conjugated to the metallic nanostructures (e.g. composite nanostructures) using similar methods as described above for the immobilization of the capture molecules to the metallic nanolayer. Such methods include, but are not limited to, EDC conjugation chemistry, sulfo-NHS coupling, hydrophobic binding or thioether chemistry. The binding partners or target analytes can be coupled to the metallic nanostructures through various chemical functionalities including thiol, amine, dithiol, acrylic phosphoramidite, azide, or alkynes.

**[0032]** In some embodiments, the metal or metals employed in the metallic nanolayer deposited on the surface can be the same as the metal or metals from which the metallic nanostructures in the detection conjugates are fabricated. For example, in one embodiment, the metallic nanolayer deposited on the surface comprises a gold film or a plurality of gold nanostructures and the detection conjugates comprise gold nanostructures. In other embodiments, the metal employed in the metallic nanolayer deposited on the surface is different from the metal or metals used to create the metallic nanostructures in the detection conjugates. For instance, in some embodiments, the metallic nanolayer deposited on the surface comprises a silver film or a plurality of silver nanostructures and the detection conjugates comprise gold nanostructures. In other embodiments, the metallic nanolayer deposited on the surface comprises a gold film or a plurality of gold nanostructures and the detection conjugates comprise silver nanostructures. In certain embodiments, the metallic nanolayer deposited on the surface comprises a gold film or a plurality of gold nanostructures and the detection conjugates comprise composite nanostructures. In related embodiments, the composite nanostructures comprise gold-coated silver nanostructures. In other particular embodiments, the metallic nanolayer deposited on the surface comprises a gold film or a plurality of gold nanostructures and the detection conjugates comprise composite nanostructures comprising gold-coated copper nanostructures. In yet other embodiments, the metallic nanolayer deposited on the surface comprises a gold film or a plurality of gold nanostructures and the detection conjugates comprise composite nanostructures comprising gold-coated magnetite nanostructures. In still other embodiments, the metallic nanolayer deposited on the surface comprises a gold film or a plurality of gold nanostructures and the detection conjugates comprise composite nanostructures comprising gold and an alkali metal or lanthanide.

**[0033]** The present disclosure also describes kits comprising the analyte detection devices as disclosed herein. In one instance, the kit comprises (i) a surface containing a metallic nanolayer on which a plurality of capture molecules are immobilized and (ii) a composition comprising a plurality of detection conjugates as described herein. In certain instances, the composition is packaged separately from the surface such that it can be brought into subsequent contact with the surface during performance of the detection methods. In some instances, the composition comprising the plurality of detection conjugates is lyophilized, for example, in the form of a pellet or bead. In related instances, the surface containing the metallic nanolayer can be a chip, disc, or a cuvette. In one particular instance, the surface containing the metallic nanolayer is a cuvette adapted for use with a centrifugal rotor. In such instances, the metallic nanolayer may be deposited on the cover, bottom and/or walls of the

cuvette.

**[0034]** In certain embodiments, all components of the analyte detection systems described herein are contained within a centrifugal rotor or disc. For instance, a rotor or disc may contain one or more reaction chambers in which the metallic nanolayer surface containing immobilized capture molecules and the plurality of detection conjugates are positioned. In one embodiment, the metallic nanolayer surface is a chip located at the bottom of the reaction chamber. In another embodiment, the metallic nanolayer is deposited directly on the floor of the reaction chamber. In still another embodiment, the metallic nanolayer surface is a bead (e.g. plastic bead) coated with the metallic nanolayer. In all such embodiments, the capture molecules are immobilized to the metallic nanolayer surfaces. In related embodiments, the plurality of detection conjugates is present in the form of a lyophilized composition, such as a lyophilized bead or pellet.

**[0035]** In alternative embodiments, capture molecules are conjugated to metallic nanostructures, which are in colloidal suspension. The plurality of detection conjugates is added to the suspension in the presence of a test sample. If the target analyte is present in the sample, complex formation will occur between the detection conjugates and the suspended nanostructures containing the capture molecules resulting in a change in optical signal (e.g., shift in peak absorbance wavelength of the suspended nanostructures).

**[0036]** Accordingly, in some instances, the kits comprise a rotor or disc having one or more reaction chambers, wherein each reaction chamber comprises (i) a lyophilized composition comprising a plurality of detection conjugates as described herein and (ii) a bead coated with a metallic nanolayer, wherein a plurality of capture molecules are immobilized to the metallic nanolayer. Such kits provide a one-step analyte detection assay whereby a test sample is contacted with the rotor or disc, and application of a centrifugal force to the rotor or disc delivers the test sample to the reaction chambers where the sample mixes with the plurality of detection conjugates and the metallic nanolayer-coated bead containing immobilized capture molecules. In embodiments in which the rotor or disc contains more than one reaction chamber, the detection conjugates and capture molecules can be selected such that a different analyte can be detected in each reaction chamber. These rotor-format detection devices can be configured in the sandwich assay format, the direct competitive format, or both if the rotors comprise multiple reaction chambers.

**[0037]** Any of the types of metallic nanolayers or metallic nanostructures discussed herein can be used with these rotor-format detection devices. In some embodiments, the metallic nanolayer coating on the bead is a gold nanolayer and the metallic nanostructures in the detection conjugates are gold nanostructures. In other embodiments, the metallic nanolayer coating on the bead is a silver nanolayer and the metallic nanostructures in the detection conjugates are gold nanostructures. In still other embodiments, the metallic nanolayer coating on the bead is a gold nanolayer and the metallic nanostructures in the detection conjugates are silver nanostructures. In one embodiment, the metallic nanolayer coating on the bead is a silver nanolayer overlaid with a gold nanolayer and the metallic nanostructures in the detection

conjugates are gold nanostructures. In certain embodiments, the metallic nanolayer coating on the bead is a gold nanolayer and the metallic nanostructures in the detection conjugates are composite nanostructures. For instance, in one embodiment, the composite nanostructures are gold-coated silver nanostructures. In another embodiment, the composite nanostructures are gold-coated copper nanostructures.

**[0038]** The kits may also include instructions for using the device to detect an analyte in a test sample, devices or tools for collecting biological samples, and/or extraction buffers for obtaining samples from solid materials, such as soil, food, and biological tissues.

**[0039]** The present invention also provides methods of detecting a target analyte in a sample. In one embodiment, the methods comprise (i) mixing a test sample with a plurality of detection conjugates as described herein; (ii) contacting the mixture with a surface containing a metallic nanolayer, wherein a plurality of capture molecules as described herein are immobilized to the metallic nanolayer; (iii) exposing the surface to a light source at a wavelength range within the ultraviolet-visible-infrared spectrum; and (iv) measuring an optical signal from the surface, wherein a change in the optical signal indicates the presence of the target analyte in the sample.

**[0040]** In some embodiments, the detection methods are sandwich assays. In such embodiments, the detection conjugates comprise metallic nanostructures coupled to binding partners that are capable of specifically binding to the target analyte if present in the sample to form analyte-detection conjugate complexes. The plurality of capture molecules, which are immobilized to the metallic nanolayer surface, are also capable of specifically binding to the target analyte if present in the sample. The metallic nanolayer is exposed to a light source and an optical signal is measured, wherein a change in the optical signal indicates the presence of analyte in the sample. By way of illustration, when a sample containing the target analyte is mixed with the plurality of detection conjugates, the target analyte binds to the binding partners in the detection conjugates to form analyte-detection conjugate complexes. These complexes in turn bind to the plurality of capture molecules immobilized to the metallic nanolayer surface through the analyte thereby bringing the metallic nanostructures in the detection conjugates in close proximity to the metallic nanolayer surface. The amount of light that is absorbed or scattered by the metallic nanolayer surface is affected by the proximity of the metallic nanostructures in the complex and thus produces an enhanced shift in the peak absorption wavelength, which indicates the presence of the target analyte in the sample.

**[0041]** In other embodiments, the detection methods are competitive assays. In such embodiments, the detection conjugates comprise metallic nanostructures coupled to the target analyte of interest. As in the sandwich assay method, the plurality of capture molecules, which are immobilized to the metallic nanolayer surface, are capable of specifically binding to the target analyte. In this type of assay, the detection conjugates will bind to the capture molecules initially. If a sample containing a target analyte is mixed with these initial complexes, the unlabeled or free target analyte in the sample will compete with the detection conjugates for binding to the capture molecules. The change in optical signal in this type of assay will result

from the displacement of the metallic nanostructures in the detection conjugates from the metallic nanolayer surface, which will proportionately reduce the wavelength shift in the peak absorption wavelength

**[0042]** A test sample can be any type of liquid sample, including biological samples or extracts prepared from environmental or food samples. In one particular embodiment, the test sample is a biological sample. Biological samples include, but are not limited to, whole blood, plasma, serum, saliva, urine, pleural effusion, sweat, bile, cerebrospinal fluid, fecal material, vaginal fluids, sperm, ocular lens fluid, mucous, synovial fluid, peritoneal fluid, amniotic fluid, biopsy tissues, saliva, and cellular lysates. The biological sample can be obtained from a human subject or animal subject suspected of having a disease condition, such as cancer, infectious diseases (e.g., viral-, bacterial-, parasitic- or fungal-infections), cardiovascular disease, metabolic disease, autoimmune disease etc. The biological sample can also be obtained from a healthy subject (e.g. human or animal) undergoing a routine medical check-up.

**[0043]** In some embodiments of the methods, the test sample is mixed with the plurality of detection conjugates and the mixture is subsequently brought into contact with the metallic nanolayer surface containing the immobilized capture molecules. In other embodiments, the test sample is contacted with the metallic nanolayer surface containing the immobilized capture molecules and the plurality of detection conjugates is subsequently added. In certain embodiments, the sample, the plurality of detection conjugates, and the metallic nanolayer surface containing the immobilized capture molecules are brought into contact at the same time. For instance, contact of the sample with both reagents simultaneously may occur in the rotor-format detection devices described above.

**[0044]** Any of the analyte detection devices described above can be used in the detection methods of the present invention. Accordingly, the various metallic nanolayer surfaces, capture molecules, and detection conjugates described herein are suitable for use in the detection methods. For instance, in some embodiments of the methods, the surface containing a metallic nanolayer is a chip, a well, a cuvette, or a bead. In certain embodiments of the methods, the surface containing a metallic nanolayer is the walls and bottom of a cuvette incorporated into or adapted for use with a centrifugal rotor. In these and other embodiments, the metallic nanolayer on the surface is a metallic film, such as a gold film. In other embodiments of the methods, the metallic nanolayer on the surface comprises a plurality of metallic nanostructures immobilized on the surface, such as gold nanostructures.

**[0045]** In certain embodiments of the detection methods, the detection conjugates comprise composite metallic nanostructures coupled to binding partners or target analytes. As described herein, composite metallic nanostructures comprise at least two noble metals or transition metals. In some embodiments of the methods, the composite metallic nanostructures comprise at least two metals selected from gold, silver, copper, platinum, palladium, ruthenium, rhodium, osmium, iridium, titanium, chromium, cadmium, zinc, iron, cobalt, and nickel. In other embodiments of the methods, the composite metallic nanostructures comprise at least two metals selected from gold, silver, copper, platinum, palladium, cadmium, iron, nickel, and zinc.

In one particular embodiment, the composite metallic nanostructures comprise gold and silver. In another embodiment, the composite metallic nanostructures comprise gold and copper. In yet another embodiment, the composite metallic nanostructures comprise silver and copper. The composite metallic nanostructures used in the methods of the invention can include a number of different geometries, such as spherical nanoparticles, pyramidal nanoparticles, hexagonal nanoparticles, nanotubes, nanoshells, nanorods, nanodots, nanoislands, nanowires, nanodisks, nanocubes, or combinations thereof.

**[0046]** In certain embodiments, the composite metallic nanostructures used in the methods of the invention are alloys of a first metal and a second metal. In some embodiments, the composite metallic nanostructures used in the methods of the invention comprise a core of a first metal and a coating of a second metal. In particular embodiments, the composite metallic nanostructures comprise a silver core and a gold coating. In other embodiments, the composite metallic nanostructures comprise a copper core and a gold coating. In another embodiment, the core is silver and the coating is copper. In some embodiments, each of the composite metallic nanostructures comprises a dielectric core (e.g. silicon dioxide, gold sulfide, titanium dioxide, silica, and polystyrene), a first coating of a first metal, and a second coating of a second metal. In one particular embodiment of the detection methods, the core is silica, the first coating (i.e. inner coating) is a silver coating, and the second coating is a gold coating (i.e. outer coating). In another embodiment, the core is silica, the first coating (i.e. inner coating) is a copper coating, and the second coating is a gold coating (i.e. outer coating).

**[0047]** The detection methods of the invention may be used to determine qualitative or quantitative amounts of a target analyte. Such methods are particularly useful for determining the approximate amount of a target analyte in a sample, which can be used *inter alia* to diagnose certain medical conditions or evaluate the efficacy of a drug therapy. In one embodiment, the quantity of a target analyte can be determined by establishing a standard curve for the particular analyte by measuring changes in optical signals from the metallic nanolayer surface as described herein for samples with a known quantity of target analyte; determining the optical signal change for a test sample; and comparing the optical signal change for the test sample to the values obtained for the standard curve. In some embodiments, determining the quantity of a complex between a first reagent and a second reagent comprises comparing the absorbance ratio and/or reaction rate from a test sample to the absorbance ratio and/or reaction rate from one sample with a known quantity of complex, thereby determining the quantity of the complex in the test sample. The quantitative values obtained from test samples may be compared to pre-determined threshold values, wherein said pre-determined threshold values are indicative of either an abnormal or normal level of the target analyte.

**[0048]** The detection methods of the present invention provide a highly sensitive technique for detecting minute quantities of a target analyte in a sample. As demonstrated by the working examples, amplification of plasmon resonance-based signals from gold nanolayer surfaces can be achieved with gold nanostructure conjugates such that nanogram quantities of target analyte can be detected in a sample. Thus, in one embodiment of the methods, the presence

of nanogram quantities of a target analyte is detected. The inventors have surprisingly found that significantly greater amplification of plasmon resonance-based signals from gold nanolayer surfaces can be achieved with composite metallic nanostructure conjugates. Use of gold-coated silver nanostructures conjugated to an analyte-specific antibody enabled the detection of pictogram quantities of the target analyte, which is a 1000-fold increase in sensitivity as compared to that obtained with gold nanostructure conjugates. See Example 3. Accordingly, in some embodiments of the methods, the presence of picogram quantities of the target analyte is detected. In other embodiments of the methods, the presence of femtogram quantities of the target analyte is detected. Greater sensitivities may be obtained by altering the composition and/or shape of the composite metallic nanostructures and/or metallic nanolayer surface.

**[0049]** When incident light is applied to metallic nanostructures, conduction band electrons in the metal oscillate collectively at the same frequency of the incident electromagnetic wave. As a result of these resonance oscillations, the nanostructures strongly absorb and scatter light at a specific wavelength range. For metallic nanostructures comprising noble or transition metals, this wavelength range is in the ultraviolet-visible-infrared spectrum depending on the particular composition of the nanostructures. Thus, light sources for applying electromagnetic energy suitable for use in the methods of the invention can include any source that may apply a wavelength range within the ultraviolet-visible spectrum or ultraviolet-visible-infrared spectrum, including arc lamps and lasers. In some embodiments, the light source may be equipped with a monochromator so that specific wavelengths of light may be applied to the metallic nanolayer surface.

**[0050]** The optical properties of the metallic nanolayers and nanostructures depend on their size, shape, and composition. For instance, solid gold nanoparticles have an absorption peak wavelength ( $\lambda_{\max}$ ) from about 515 nm to about 560 nm depending on particle size. Gold spherical nanoparticles having a 30 nm diameter maximally absorb at about 520 nm with  $\lambda_{\max}$  shifting to longer wavelengths as particle diameter increases. Silver and copper particles have a  $\lambda_{\max}$  in the ultra-violet/blue or red region (e.g., from about 350 nm to about 500 nm) with increasing particle diameter causing a shift in  $\lambda_{\max}$  to longer wavelengths. Metallic nanorods have a transverse  $\lambda_{\max1}$  and a longitudinal  $\lambda_{\max2}$ . Alloys of different metals typically exhibit absorption peaks in an intermediate range between the absorption peaks of the comprising metals. For example, nanostructures comprising a 50/50 alloy of gold and silver exhibit a  $\lambda_{\max}$  of about 470 nm with increasing amounts of gold causing a shift in the absorption peak to longer wavelengths. The sensitivity of the LSPR signals to changes in the local medium refractive index can be modified by changing the shape or geometry of the nanostructures. For instance, nonspherical particles (e.g. nanoprisms, nanorods, nanoshells, etc.) have increased LSPR sensitivities as compared to spheres. In some embodiments, the optical properties (e.g. absorption/scattering at particular wavelengths) are tailored to a particular application by varying the size, shape, or composition of the metallic nanolayer deposited on the surface or the metallic nanostructures employed in the detection conjugates.



**[0051]** The interaction between the incident light and the metallic nanolayer surface can be monitored as reflected light or transmitted light. The amount of the incident light that is absorbed or scattered can be measured as an absorption spectrum in a reflection mode or the absorption spectrum in a transmission mode. In some embodiments, the optical signal measured from the metallic nanolayer can be an optical reflection, an absorbance spectrum, a scattering spectrum, and/or an emission spectrum.

**[0052]** The plasmon coupling between the metallic nanolayer and the metallic nanostructures in the detection conjugates resulting from complex formation between the binding partners, target analyte, and capture molecules produces a change in the localized surface plasmon resonance spectrum of the metallic nanolayer. For instance, such changes can include an increased optical extinction, an increased optical reflection, and/or increased scattering and/or emission signal. In some embodiments, the change in optical signal indicative of the presence of the target analyte in the sample includes a shift, increase or decrease in optical scattering or a combination of these features. In certain embodiments, the change in optical signal indicative of the presence of the target analyte in the sample is a spectral peak wavelength shift. In one embodiment, the wavelength shift in the optical spectral peak may be a red shift (e.g., a shift to a longer wavelength) within a 200 nm to 1200 nm spectral window. In another embodiment, the wavelength shift in the optical spectral peak may be a blue shift (e.g., a shift to a shorter wavelength) within a 200 nm to 1200 nm spectral window. The changes in optical signals can be measured at a particular time point following a set reaction period. Additionally or alternatively, changes in the optical signal over the reaction period (e.g. rate determinations) may be measured. Both types of measurements can be used for either qualitative or quantitative analysis of a target analyte.

**[0053]** Various means for measuring optical signals at different wavelengths and acquiring extinction, scattering, or emission spectra are known in the art. Any spectrophotometric or photometric instruments are suitable for use in the disclosed methods. Some non-limiting examples include plate readers, Cobas Fara analyzers, and Piccolo xpress® and Vetscan analyzers (Abaxis, Inc., Union City, CA), optic fiber readers (e.g., LightPath™ S4 (LamdaGen, Menlo Park, CA)), SPR instruments (e.g., Biacore instruments available from GE Healthcare), centrifugal analyzers from Olympus, Hitachi etc.

**[0054]** The present invention also includes an assay complex comprising (i) a detection conjugate that comprises a composite metallic nanostructure coupled to a binding partner, (ii) a target analyte, and (iii) a metallic nanolayer-coated bead on which a capture molecule is immobilized, wherein the binding partner in the detection conjugate is bound to a first epitope on the target analyte and the capture molecule is bound to a second epitope on the target analyte, thereby forming a complex comprising the detection conjugate, target analyte, and the capture molecule. In some embodiments, the assay complex is contained within a cuvette adapted for use with a centrifugal rotor. In other embodiments, the assay complex is contained within a reaction chamber in a centrifugal rotor or disc.

**[0055]** The binding partner and capture molecule in the assay complex can be any type of

molecule described above, including haptens and other small molecules, drugs, hormones, biological macromolecules such as antibodies or fragments thereof (e.g., Fv, Fab, (Fab)<sub>2</sub>, single chain, CDR etc.), antigens, receptors, ligands, polynucleotides, aptamers, polypeptides, polysaccharides, lipopolysaccharides, glycopeptides, lipoproteins, or nucleoproteins. In one embodiment, the binding partner is an antibody and the capture molecule is a different antibody.

**[0056]** The metallic nanolayer and composite metallic nanostructures are described in detail above. In one embodiment, the metallic nanolayer coating the bead (e.g. plastic or glass bead) is a gold nanolayer. In another embodiment, the metallic nanolayer coating the bead is a silver nanolayer. The bead is preferably less than 0.5 cm, but greater than 0.1 mm. In certain embodiments, the composite metallic nanostructures are gold-coated silver nanostructures. In other embodiments, the composite metallic nanostructures are gold-coated copper nanostructures. In still other embodiments, the metallic nanostructures comprise gold doped with silver, copper ions or both of these ions.

**[0057]** Any type of target analyte can be detected using the methods, devices, and assay complexes of the present invention, particularly those that are significant in the diagnoses of diseases. A target analyte can include, but is not limited to, a protein, enzyme, antigen, antibody, peptide, nucleic acid (RNA, DNA, mRNA, miRNA), hormone, glycoprotein, polysaccharide, toxin, virus, virus particle, drug molecule, hapten, or chemical. In some embodiments, the target analyte is a marker or antigen associated with an infectious disease in humans and/or animals. In other embodiments, the target analyte is a marker or antigen associated with a particular physiological state or pathological condition.

**[0058]** In certain embodiments, the target analyte is a pathogenic antigen or antibody to a pathogenic antigen. For instance, the pathogenic antigen can be a viral antigen (e.g., feline leukemia virus, canine parvovirus, foot and mouth virus, influenza virus, hepatitis a, b, c virus, HIV virus, human papilloma virus, Epstein Barr virus, rabies virus, etc.), a bacterial antigen (e.g., *Ehrlichia*, *Borrelia*, *Anaplasma*, *Anthrax*, *Salmonella*, *Bacillus*, etc.), a fungal antigen, or parasitic antigen (e.g., canine heartworm, *Giardia lamblia*, plasmodium falciparum, African trypanosomiasis, *Trypanosoma brucei*, etc.). In other embodiments, the target analyte is a disease-related antigen or antibody to a disease-related antigen. Disease-related antigens include, but are not limited to, cancer-related antigens or markers (e.g., PSA, AFP, CA125, CA15-3, CA19-9, CEA, NY-ESO-1, MUC1, GM3, GD2, ERBB2, etc.), cardiovascular disease-related antigens or markers (e.g., troponin, C-reactive protein, brain natriuretic peptide, CKMB, fatty acid binding protein, etc.), metabolic-related antigens or markers (e.g., thyroid stimulating hormone, thyroxine, leptin, insulin), or autoimmune disease-related antigens or markers (e.g., auto-antibodies). In certain embodiments, the target analyte is an inflammatory antigen or marker (e.g., C-reactive protein, MRP14, MRP8, 25F9, etc.). In other embodiments, the target analyte is a pregnancy-related antigen or marker (e.g., a fetal antigen, human chorionic gonadotropin).

**[0059]** The present disclosure also describes a method for preparing composite metallic

nanostructures. In one instance, the method comprises preparing a first solution comprising a mixture of a polymer and chloroauric acid, preparing a second solution comprising silver or copper nanostructures, and incubating the first solution with the second solution for a period of time, wherein the resulting mixture comprises gold-coated silver nanostructures or gold-coated copper nanostructures. The resulting mixture preferably has a peak absorbance of about 515 nm to about 670 nm, or about 520 nm to about 560 nm. In one embodiment, the resulting mixture has a peak absorbance of about 530 nm.

**[0060]** The polymer used in the preparation of the first solution can be any one of polyvinylpyrrolidone, polyvinyl alcohol, polyacrylate, polyethylene glycol, polyethyleneimine, polyaspartic acid, polyglutamic acid, various gums, gelatin or mixed polymers comprising any of the foregoing. In one particular instance, the polymer is polyvinylpyrrolidone. Different types of coated nanostructures can be obtained by varying the molecular weight of the polymer. Suitable molecular weight ranges of the polymer include from about 5,000 Daltons to about 150,000 Daltons, about 10,000 Daltons to about 100,000 Daltons, from about 20,000 Daltons to about 80,000 Daltons. In some instances, the polymer has a molecular weight less than 50,000 Daltons. In other instances, the polymer has a molecular weight less than 20,000 Daltons. In certain instances, the polymer has a molecular weight of about 10,000 Daltons.

**[0061]** The characteristics of the gold coating can be controlled by adjusting the concentration ratio of polymer to chloroauric acid. For instance, the concentration ratio of polymer to chloroauric acid is from about 100:1 to about 1:100, from about 2:1 to about 5:1, or from about 1.5:1 to about 8:1. In some instances, the concentration ratio of polymer to chloroauric acid is 1:1. Suitable concentrations of polymer include, but are not limited to, about 0.1 % to about 20% wt/wet in water or ethanol. Suitable concentrations of chloroauric acid include, but are not limited to, about 0.001 M to about 1.0 M, about 0.010 M to about 0.500 M, and about 0.050 M to about 0.100 M.

**[0062]** The coating efficiency and thickness can also be affected by the pH and halide content of the coating solution (i.e. first solution). In certain instances, the pH of the solution is kept in a range from about 3 to about 14. The halide content of the solution is, in some instances, less than 150 mM. In other instances, the halide content of the solution is in the range of about 0 to about 50 mM.

**[0063]** Methods of preparing solutions of silver and copper nanostructures are known to those of skill in the art. For instance, the second solution comprising silver or copper nanostructures can be prepared by any of the methods described in U.S. Patent Publication No. 2012/0101007, U.S. Patent Publication No. 2014/0105982, or U.S. Patent Publication No. 2013/0230717. In one instance, the second solution comprising silver or copper nanostructures is prepared by mixing a silver or copper source with a reducing agent. A suitable silver source includes a silver salt, such as silver nitrate. Suitable copper sources include copper (II) sulfate, copper (II) chloride, copper (II) hydroxide and copper (II) nitrate, copper (II) acetate and copper (II) trifluoroacetate. Reducing agents that can be reacted with the silver or copper sources to form the nanostructures can include glucose, ascorbic acid,

sodium borohydride, and alkaline solutions (e.g. pH greater than 7.5) of polymers such as PVP. In certain instances, the reducing agent is ascorbic acid. The desired shape and optical spectral peak of the silver nanostructures or copper nanostructures can be attained by adjusting the ratios or concentrations of reactants as known to those of ordinary skill in the art. By way of example only, high concentrations of the reducing agent can result in pentagonal- and bipyramidal-shaped nanostructures, whereas low concentrations of the reducing agent can result in elongated nanowires or tubes. Depending on the particular shapes of the nanostructures, the second solution comprising silver or copper nanostructures may have a peak absorbance from about 550 nm to about 1000 nm, from about 600 nm to about 700 nm, from about 630 nm to about 680 nm, from about 750 nm to about 850 nm, from about 900 nm to about 940 nm, from about 580 nm to about 620 nm, or from about 550 nm to about 750 nm. In certain instances, the second solution comprising silver nanostructures has a peak absorbance of about 600 nm (i.e. 595 nm to 605 nm, inclusive). In some instances, the second solution comprising copper nanostructures has a peak absorbance of about 585 nm (i.e. 580 nm to 590 nm, inclusive). In some instances, the peak absorbance of a solution comprising copper nanostructures is greater (i.e. red-shifted) than the peak absorbance of a solution comprising silver nanostructures of a similar size and shape.

**[0064]** In some instances, the incubation period of the first solution with second solution is at least 12 hours. In other instances, the incubation period of the first solution with second solution is greater than 24 hours, preferably greater than 48 hours, more preferably at least 72 hours. Changes in the peak absorbance of the reaction mixture can be monitored during the incubation period to adjust the incubation time accordingly. For example, shifts of the peak absorbance to shorter wavelengths, for instance in the 520 nm to 550 nm region, can indicate that the gold-coated nanostructures have stabilized. In certain instances, stability of the resulting nanostructures to sodium chloride (e.g., 0.25-1M) is used to indicate a proper coating of the nanostructures.

**[0065]** The present disclosure further describes methods of synthesizing nanostructures having optical densities greater than about 50/mL. In one instance, the methods comprise mixing a polymer as described herein with chloroauric acid, stirring the mixture at a set temperature for a first period of time, adding ascorbic acid to the mixture, and incubating the mixture for a second period of time. The size and shape of the nanostructures is dictated by the concentration ratio of polymer to chloroauric acid and the temperature and time of incubation. The concentrations of polymer and chloroauric acid can be in the ranges described above. The temperature can be adjusted based on the size and shape of the nanostructures desired, but may be in the range of about 4 °C to about 100 °C. Similarly, the incubation period (i.e. first period of time) can be adjusted based on the desired properties of the nanostructures, but may range from about 15 minutes to one day.

**[0066]** In some instances, about 0.1 to 1 part of ascorbic acid (e.g. about 1 to 5 M) is added to the mixture following the first incubation period. The second incubation period following addition of the ascorbic acid may be from about 1 to about 24 hours. Without being bound by theory, addition of ascorbic acid provides a substantial increase in the quantity of

nanostructures produced.

**[0067]** In certain instances, the methods further comprise adding or doping the mixture with about 1 to about 100 parts of gold chloride (e.g. about 0.001 M to 1M) or silver nitrate (e.g. about 0.001 M to 1M) or other metal (e.g. noble metal, transition metal, alkali metal, or lanthanide). This doping step can further increase the resonance intensity of the resulting nanostructures. In some instances, the gold chloride, silver nitrate, or other metal is added to the mixture before ascorbic acid is added to the reaction. In other instances, the gold chloride, silver nitrate, or other metal is added to the mixture following the addition of ascorbic acid. The order of addition of the metal and ascorbic acid may be adjusted to tailor the resulting nanostructures to a desired shape and size.

**[0068]** This invention is further illustrated by the following additional examples that should not be construed as limiting. Those of skill in the art should, in light of the present disclosure, appreciate that many changes can be made to the specific embodiments which are disclosed and still obtain a like or similar result without departing from the scope of the invention.

**[0069]** The invention is defined by the appended claims.

## EXAMPLES

### Example 1. Amplification of LSPR signals with Gold Nanoparticle-Conjugated Analyte

**[0070]** An analyte detection system was prepared by providing a plastic chip on which was deposited a gold nanolayer film. Human IgG proteins (100 µg/ml) were immobilized to the gold nanolayer film to create the sensor surface. A control sensor was constructed by immobilizing bovine serum albumin to the gold nanolayer film. The two types of sensor surfaces were positioned within an instrument equipped with light emitting and light collecting fibers that shine light on the gold nanolayer surface and collect the light reflected back from the surface.

**[0071]** A sample containing free protein A (10 µg/ml) was contacted with the two types of sensor surfaces and the changes in the reflectance spectra were measured. As shown in Figure 1, introduction of free protein A to the sensor containing immobilized human IgG does not produce a significant visible change in the reflectance spectrum of the gold nanolayer film as measured by a shift in the peak wavelength.

**[0072]** The sensor surfaces were regenerated by treatment with 1 mM hydrochloric acid and sample containing protein A conjugated to colloidal gold nanoparticles (CGC) at two different concentrations (3.5 µg/ml and 0.175 µg/ml) were contacted with the sensor surfaces. The change in the reflectance spectrum of the gold nanolayer surface was accentuated when the protein A (i.e. target analyte) was conjugated to colloidal gold nanoparticles. Specifically, 3.5

$\mu\text{g/ml}$  of protein A-CGC produced a greater shift in peak wavelength than  $10 \mu\text{g/ml}$  of unlabeled protein A. See Figure 1, sensor 2. The amplification of the plasmon resonance signal was sufficiently great to enable detection of nanogram concentrations of protein A-CGC. See Figure 1, sensor 3. The changes in the reflectance spectrum of the BSA sensor represent non-specific binding of the protein A molecules to the sensor surface and are significantly smaller than the changes induced by specific binding of the protein A molecules to the immobilized IgG molecules.

**[0073]** The results of this initial experiment show that considerable amplification of the changes in localized surface plasmon resonance signals induced by binding events at a metallic nanolayer surface can be achieved by coupling the target analyte to colloidal gold nanoparticles. Nearly a 60-fold improvement in sensitivity is observed with nanogram quantities of analyte being detected.

#### **Example 2. Amplification of LSPR signals in a Sandwich Assay**

**[0074]** This example describes a series of experiments designed to evaluate whether amplification of localized surface plasmon resonance signals with gold nanoparticle conjugates could also be achieved in a sandwich assay format in which the target analyte is not directly conjugated to the gold nanoparticles. A gold nanolayer chip surface was prepared as described in Example 1. The C7 antibody against C-reactive protein (CRP) ( $100 \mu\text{g/ml}$ ) was immobilized to the gold nanolayer film deposited on the chip surface to create the anti-CRP sensor. The C6 antibody, which recognizes a distinct, non-overlapping epitope of CRP than the C7 antibody, was conjugated to colloidal gold nanoparticles (C6-CGC) for some experiments or used in an unlabeled form for other experiments.

**[0075]** In a first series of experiments, a sample containing one of three different concentrations of CRP ( $1 \text{ ng/ml}$ ,  $10 \text{ ng/ml}$ , or  $100 \text{ ng/ml}$ ) was incubated with the anti-CRP sensor for 15 to 20 minutes and changes in the reflectance spectrum of the gold nanolayer were monitored. As shown in Figure 2, very minimum peak shift was observed upon binding of the CRP to the immobilized C7 anti-CRP antibody on the sensor surface. Subsequent exposure of the sensor surface to unlabeled C6 anti-CRP antibody ( $1 \mu\text{g/ml}$ ) did not result in significant further peak shifts. See Figure 2. Similarly, subsequent exposure of the sensor surface to  $3 \mu\text{g/ml}$  of C6-CGC did not produce any further changes in the reflectance spectrum, indicating that the bound CRP molecules were likely saturated with unlabeled C6 antibody. See Figure 2.

**[0076]** In a second series of experiments, a sample containing one of three different concentrations of CRP ( $1 \text{ ng/ml}$ ,  $10 \text{ ng/ml}$ , or  $100 \text{ ng/ml}$ ) was incubated with the anti-CRP sensor for 15 to 20 minutes. Two different concentrations of C6-CGC ( $1 \mu\text{g/ml}$  and  $3 \mu\text{g/ml}$ ) were subsequently introduced and changes in the reflectance spectrum were measured (Figures 3 and 4A). The results show that conjugation of the C6 anti-CRP antibody to gold nanoparticles amplifies the peak wavelength shift as compared to unlabeled C6 antibody.

Increasing concentrations of C6-CGC produce a dose-dependent shift in peak wavelength. However, the signal difference between 1 ng/ml and 10 ng/ml was small (Figure 4B).

**[0077]** In a third series of experiments, the effect of analyte incubation time on signal development was evaluated. The anti-CRP sensor was contacted with sample containing 0 ng/ml, 10 ng/ml, or 100 ng/ml CRP and 3  $\mu$ g/ml C6-CGC was immediately introduced without any analyte incubation time or a wash. As shown in Figures 5 and 6, shorter analyte incubation time results in smaller peak wavelength shifts.

**[0078]** The results of these three sets of experiments show that amplification of LSPR signals can be achieved with gold nanoparticle conjugates in a sandwich assay format. An enhanced signal shift is observed when the detector antibody is labeled with colloidal gold particles as compared to unlabeled antibody, thereby allowing for detection of nanogram concentrations of analyte.

### **Example 3. Enhanced Signal Amplification with Gold-Coated Silver Nanostructures**

**[0079]** To examine whether varying the type of metal used to label the binding partners affected the amplification of LSPR signals, composite metal nanostructures were prepared. Specifically, gold-coated silver nanostructures were prepared as follows. Silver nanostructures were prepared by adding 50.0 mL of deionized H<sub>2</sub>O, 500.0  $\mu$ L Trisodium Citrate (75mM), 200  $\mu$ L AgNO<sub>3</sub> (200mM), and 500.0  $\mu$ L H<sub>2</sub>O<sub>2</sub> (27%) while stirring vigorously at room temperature. A 500  $\mu$ L aliquot of NaBH<sub>4</sub> (200mM) was then rapidly injected into the aqueous solution causing a color change to light yellow. Over a period of several minutes the color continued to change from dark yellow to red to purple and finally stabilizing at blue. The UV/Vis spectra determined the peak absorbance of the solution to be at 604.5 nm.

**[0080]** A gold coating was added to the silver nanostructures by adding 5.0 mL of the blue solution to a mixture of 50  $\mu$ L of Polyvinylpyrrolidone (PVP MW  $\approx$  10,000 20% in ethanol) and 50  $\mu$ L of HAuCl<sub>4</sub> (20mM). After 72 hours of incubation time the sample became a dark red color and had a peak absorbance at 534.5 nm. The nanoparticles were washed twice by centrifugation at 20,000 rpm for 20 minutes and resuspended in 2.0 mL of deionized H<sub>2</sub>O. The solution had a deep red color, absorption peak at 530.3 nm, and a total absorbance of 15.0 OD units.

**[0081]** Conjugation of gold-coated silver nanostructures (Au@AgNPs) to the C6 anti-CRP antibody was performed by adding 600.0  $\mu$ L of Au@AgNPs and 20.0  $\mu$ L C6 anti-CRP antibody (8.0 mg/mL) to 880.0  $\mu$ L deionized H<sub>2</sub>O bringing the final antibody concentration to 17.8  $\mu$ g/mL/OD. After a 2 hour incubation period at 4 °C, the sample was centrifuged at 30,000g for 20 minutes and resuspended in 1.5 mL of a blocking solution containing BSA (10 mg/mL) in PBS. The Au@AgNPs conjugated to anti-CRP C6 antibody were stored at 4 °C until further use.

**[0082]** The anti-CRP gold nanolayer sensor was prepared as described in Example 2 and had peak absorption at 530nm. A control sensor containing the gold nanolayer without any immobilized antibody was also prepared. The sensors were equilibrated with 100  $\mu$ L PBS.

**[0083]** 100  $\mu$ L C6 anti-CRP antibody conjugated to Au@AgNPs diluted to 1.5 OD in PBS was premixed for 1 minute with 1, 10, or 500 pg/mL CRP antigen. The mixture was then brought into contact with the anti-CRP or control sensor surface and changes in the reflectance spectrum of the gold nanolayer surface were measured. The results show that the gold-coated silver nanostructures enhanced the peak wavelength shift induced by binding of the CRP-antibody complex to the sensor surface (Figure 7). Detection of 1 pg/mL of CRP antigen was possible with the gold-coated silver nanostructures, which is a 1000-fold improvement in sensitivity as compared to that obtained with gold nanoparticles. At higher concentrations of antigen the binding sites are saturated and no further shifts occur.

**[0084]** The results of this experiment demonstrate the significantly enhanced amplification of LSPR signals from a metallic nanolayer surface achieved when composite nanostructures, such as gold-coated silver nanostructures, are used to label analyte binding partners.

#### **Example 4. Synthesis of High Optical Density Nanostructures**

**[0085]** Gold nanoparticles were prepared by mixing the following reagents in a final volume of 1 ml in the indicated order: 0.1 ml of 1% PVP-10 (1% wt/wt), 0.2 ml of 0.1M gold chloride, 0.1 ml of 5N NaOH, 0.4 ml of water and 0.2 ml of 1M ascorbic acid. The reaction mixture was mixed after each addition. The spectroscopic measurements indicated that the reaction was mostly complete after 24 hours at room temperature. This protocol yielded spherical gold nanoparticles exhibiting the LSPR peak around 535 nm and the corresponding optical density of about 80 per ml. Layering with additional gold or silver was done by adding silver nitrate or gold chloride to the preformed gold nanoparticles. Excess reagents were removed by centrifugation at 30,000 g for 1-2 hours.

**[0086]** In a separate reaction, 0.05 ml of 20% PVP (wt/wt) was mixed with 0.25 ml water, 0.1 ml of 5N NaOH, 0.1 ml of 1 M sodium citrate, 0.5 ml of 0.1M gold chloride and 1 ml of 1M ascorbic acid. This protocol resulted in immediate formation of colloidal gold particles at an OD of about 90/ml with LSPR peak at ~525 nm. A linear correspondence was observed between final OD and the concentration of gold between 2.5 mM gold and 25 mM gold in the final reaction mixture.

**[0087]** It is understood that the disclosed invention is not limited to the particular methodology, protocols and materials described as these can vary. It is also understood that the terminology used herein is for the purposes of describing particular embodiments only and is not intended to limit the scope of the present invention which will be limited only by the appended claims.



[0088] Those skilled in the art will recognize, or be able to ascertain using no more than routine experimentation, many equivalents to the specific embodiments of the invention described herein. Such equivalents are intended to be encompassed by the following claims.

## REFERENCES CITED IN THE DESCRIPTION

### Cited references

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#### Patent documents cited in the description

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- [US20120101007A](#) [0063]
- [US20140105982A](#) [0063]
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- ATANASOV et al. Journal of Physics: Conference Series, 2014, vol. 514, [0017]
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- GUPTA et al. J. Appl. Phys., 2002, vol. 92, 5264-5271 [0017]

**P a t e n t k r a v**

**1.** Analytdetekteringsindretning til registrering baseret på lokaliseret overflade-plasmonresonans, hvilken indretning omfatter:

- 5 en flerhed af detekteringskonjugater, hvor konjugaterne omfatter sammensatte metalliske nanostrukturer koblet til bindingspartnere, der er antistoffer, der kan binde specifikt til en første epitop af en målanalyt, hvor de sammensatte metalliske nanostrukturer omfatter en guldcoating og en sølvkerne;  
en overflade indeholdende et metallisk nanolag; og  
10 en flerhed af indfangningsmolekyler, der er immobiliseret på det metalliske nanolag, hvor indfangningsmolekylerne er antistoffer, der kan binde specifikt til en anden epitop af målanalytten.

**2.** Analytdetekteringsindretning ifølge krav 1, hvor de sammensatte metalliske nanostrukturer er sfæriske nanopartikler og har en diameter på 5 nm til 200 nm eller 10 nm til 100 nm; eller hvor de sammensatte metalliske nanostrukturer er nanoplader med en kantlængde på 10 nm til 800 nm og en tykkelse på 1 nm til 100 nm.

**3.** Analytdetekteringsindretning ifølge krav 1, hvor flerheden af detekteringskonjugater er i form af en lyofiliseret pellet eller perle.

**4.** Analytdetekteringsindretning ifølge krav 1, hvor overfladeen er en chip, en brønd, en perle eller en væg, et låg og/eller en bund af en kuvette.

**5.** Analytdetekteringsindretning ifølge krav 1, hvor det metalliske nanolag er en metallisk film; eller  
hvor det metalliske nanolag omfatter en flerhed af metalliske nanostrukturer, der er immobiliseret på overfladen.

**6.** Analytdetekteringsindretning ifølge krav 1, hvor de sammensatte nanostrukturer har en geometri, der er udvalgt blandt sfæriske nanopartikler, pyramide-

formede nanopartikler, sekskantede nanopartikler, nanoskaller, nanorør, nanostave, nanopunkter, nanoøer, nanotråde eller kombinationer deraf.

5       **7.** Analytdetekteringsindretning ifølge krav 1, hvor målanalytten er en markør eller et antigen, der er forbundet med en infektionssygdom, fysiologisk tilstand eller patologisk tilstand, eller hvor målanalytten er hunde-hjerteorm, feline leukæmivirus, hunde-parvovirus, C-reactivt protein, *Giardia lamblia*, *Ehrlichia*-antigen eller -antistof, *Borrelia*-antigen eller -antistof, *Anaplasma*-antigen eller -antistof, et cancerantigen, et hjertemarkør-antigen, thyreoideastimulerende  
10       hormon, thyroxin, troponin eller hjernenatriuretisk peptid.

15       **8.** Analytdetekteringsindretning ifølge krav 1, hvor de sammensatte metalliske nanostrukturer har et absorptionspeak på 515 nm til 670 nm eller 520 nm til 560 nm.

20       **9.** Analytdetekteringsindretning ifølge krav 1, hvor indretningen giver signifikant større forbedring i førsomhed sammenlignet med en anden indretning, der er den samme som indretningen, bortset fra at de sammensatte metalliske nanostrukturer består af guld.

25       **10.** Fremgangsmåde til detektering af en målanalyt i en prøve baseret på lokaliseret overfladeplasmonresonans under anvendelse af analytdetekteringsindretningen ifølge et hvilket som helst af de foregående krav, hvilken fremgangsmåde omfatter:

30       at blande prøven med flerheden af detekteringskonjugater, hvor konjugaterne omfatter sammensatte metalliske nanostrukturer koblet til bindingspartnere, der er antistoffer, der kan binde specifikt til den første epitop af en målanalyt, hvis de er til stede i prøven, til dannelse af analytdetekteringskonjugatkomplekser, hvor de sammensatte metalliske nanostrukturer omfatter en guldcoating og en sølvkerne;

at bringe blandingen i kontakt med overfladen indeholdende det metalliske nanolag med flerheden af indfangningsmolekyler, der er immobiliseret derpå, hvor indfangningsmolekylerne er antistoffer, der kan binde specifikt til den anden epitop af målanalytten;

at udsætte overfladen for en lyskilde ved et bølglængdeområde inden for ultraviolet-synlig-infrarød-spektret; og  
 at måle et optisk signal fra overfladen, hvor en ændring i det optiske signal indikerer tilstedeværelse af analytten i prøven.

5

**11** Fremgangsmåde ifølge krav 10, hvor

a) det optiske signal er reflektans, et absorptionsspektrum, et spredningsspektrum eller et emissionsspektrum;

b) ændringen i det optiske signal omfatter et spektralt peakbølglængdeskift;

10 c) tilstedeværelsen af nanogram-, picogram- eller femtogram-mængder af målanalytten detekteres; eller

d) overfladen er væggene og bunden af en kuvette, der er inkorporeret i en centrifugal rotor.

15

**12.** Kompleks til udførelse af et assay til registrering baseret på lokaliseret overfladeplasmonresonans, omfattende:

et detekteringskonjugat omfattende en sammensat metallisk nanostruktur koblet til en bindingspartner, der er et antistof, hvor den sammensatte metalliske nanostruktur omfatter en guldcoating og en sølvkerne;

20

en målanalyt med en første epitop og en anden epitop; og

en perle coatet med et metallisk nanolag, hvorpå der er immobiliseret et indfangningsmolekyle, hvor: (a) indfangningsmolekylet er et antistof, der er et andet antistof end bindingspartnerantistoffet, og (b) bindingspartneren i detekteringskonjugatet er bundet til den første epitop på målanalytten, og indfangningsmolekylet er bundet til den anden epitop på målanalytten, hvorved der dannes det kompleks, der omfatter detekteringsmolekylet, målanalytten og indfangningsmolekylet.

25

**13.** Kompleks ifølge krav 12, hvor

30

a) det metalliske nanolag er en metallisk film; eller

b) det metalliske nanolag omfatter en flerhed af metalliske nanostrukturer, der er immobiliseret på perlen.

**14.** Analytdetekteringsindretning ifølge krav 5 eller kompleks ifølge krav 13, hvor den metalliske film omfatter guld, sølv, kobber, platin, palladium, cadmium, zink eller et komposit deraf.

5      **15.** Analytdetekteringsindretning ifølge krav 5 eller kompleks ifølge krav 13, hvor flerheden af de metalliske nanostrukturer omfatter guld, sølv, kobber, platin, palladium, cadmium, zink eller et komposit deraf.

10      **16.** Analytdetekteringsindretning, fremgangsmåde eller kompleks ifølge krav 1, 10 eller 12, hvor indretningen, fremgangsmåden eller assayet af komplekset har en følsomhed til detektering af picogram-mængder af målanalytten.

## DRAWINGS

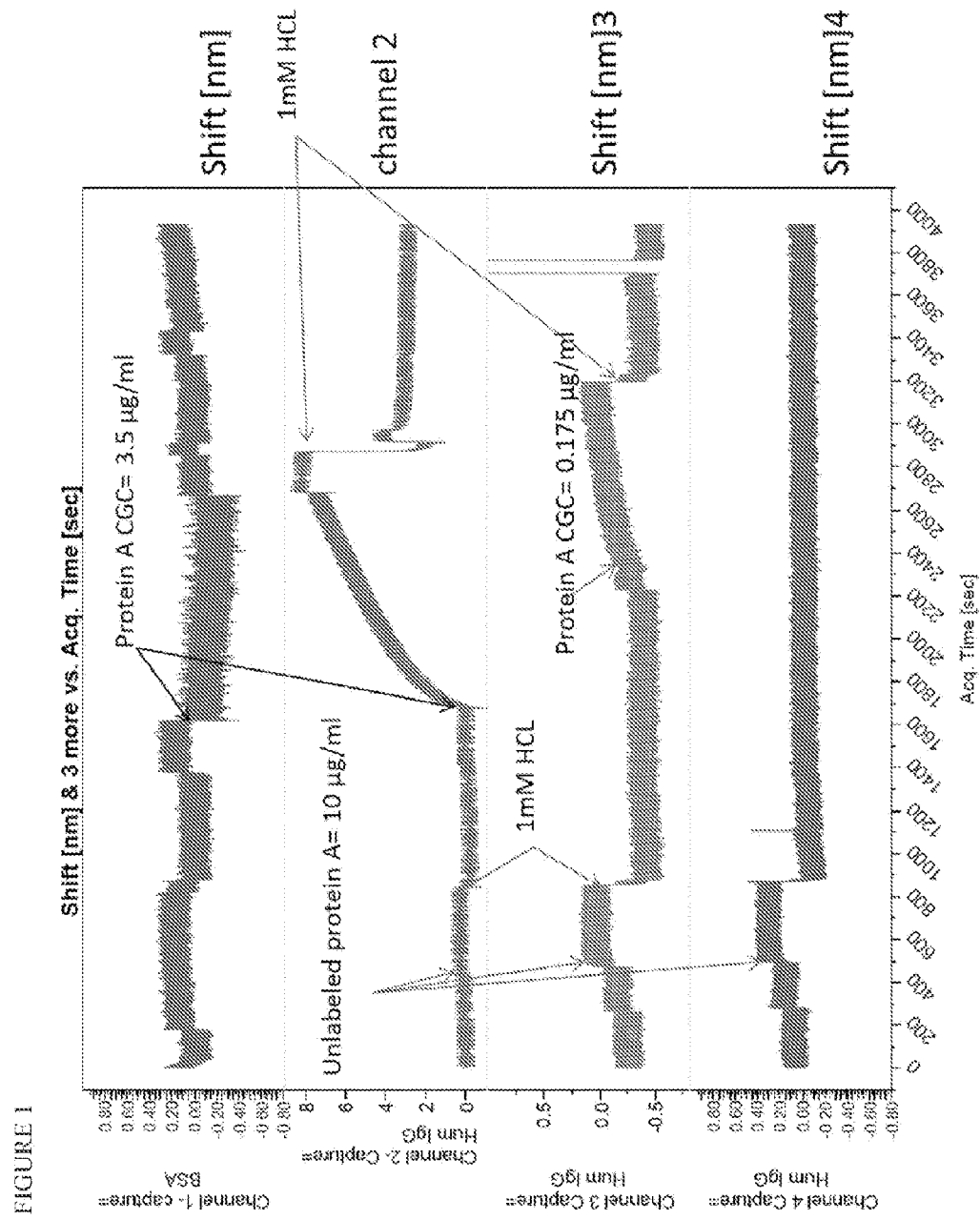


FIGURE 2

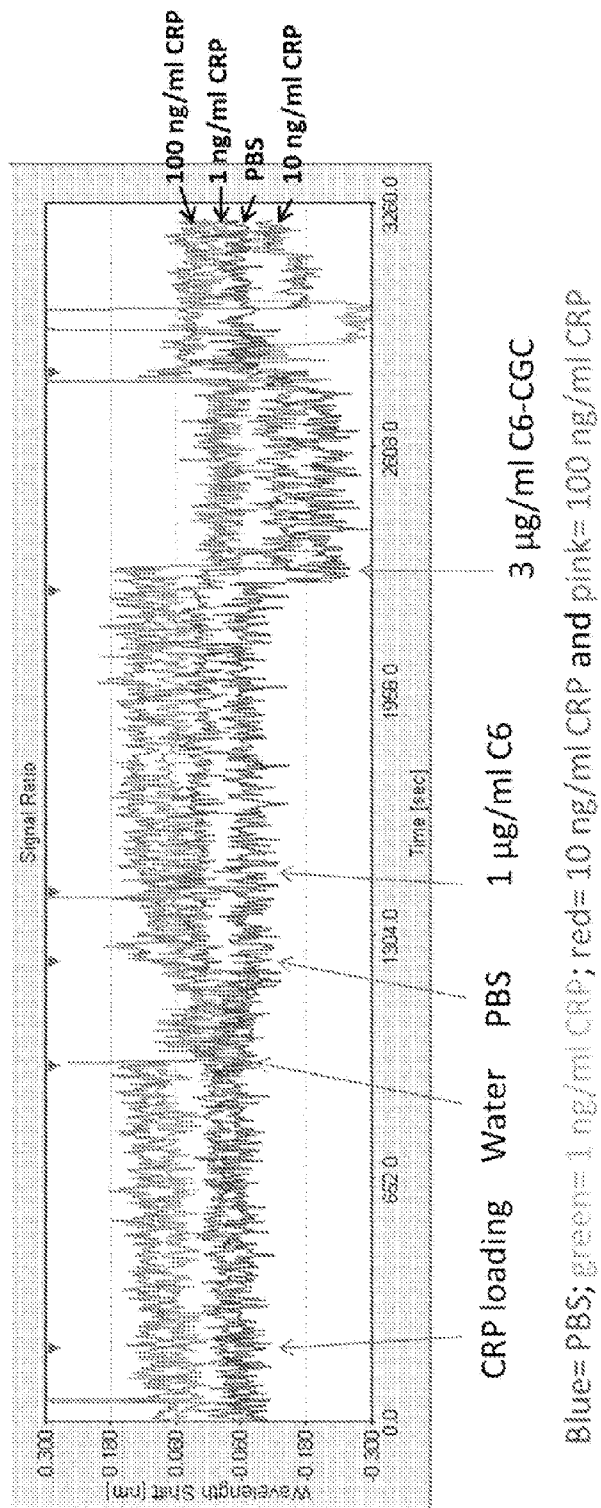


FIGURE 3

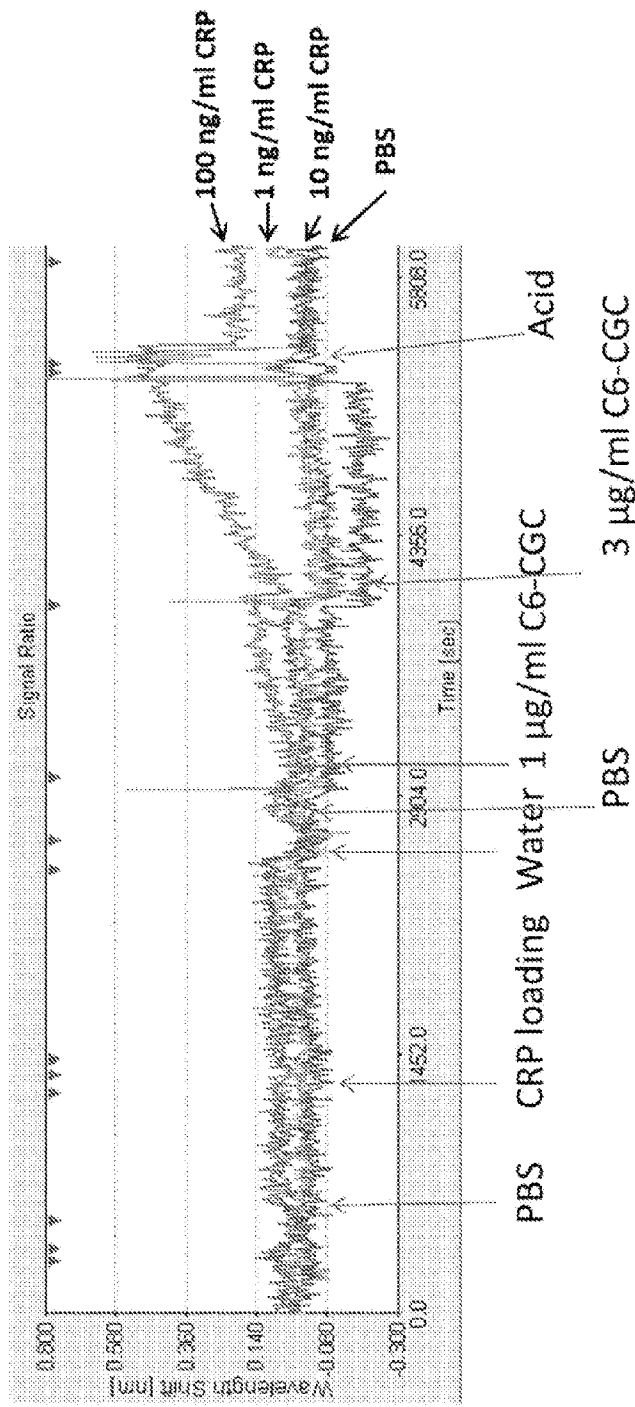




FIGURE 4

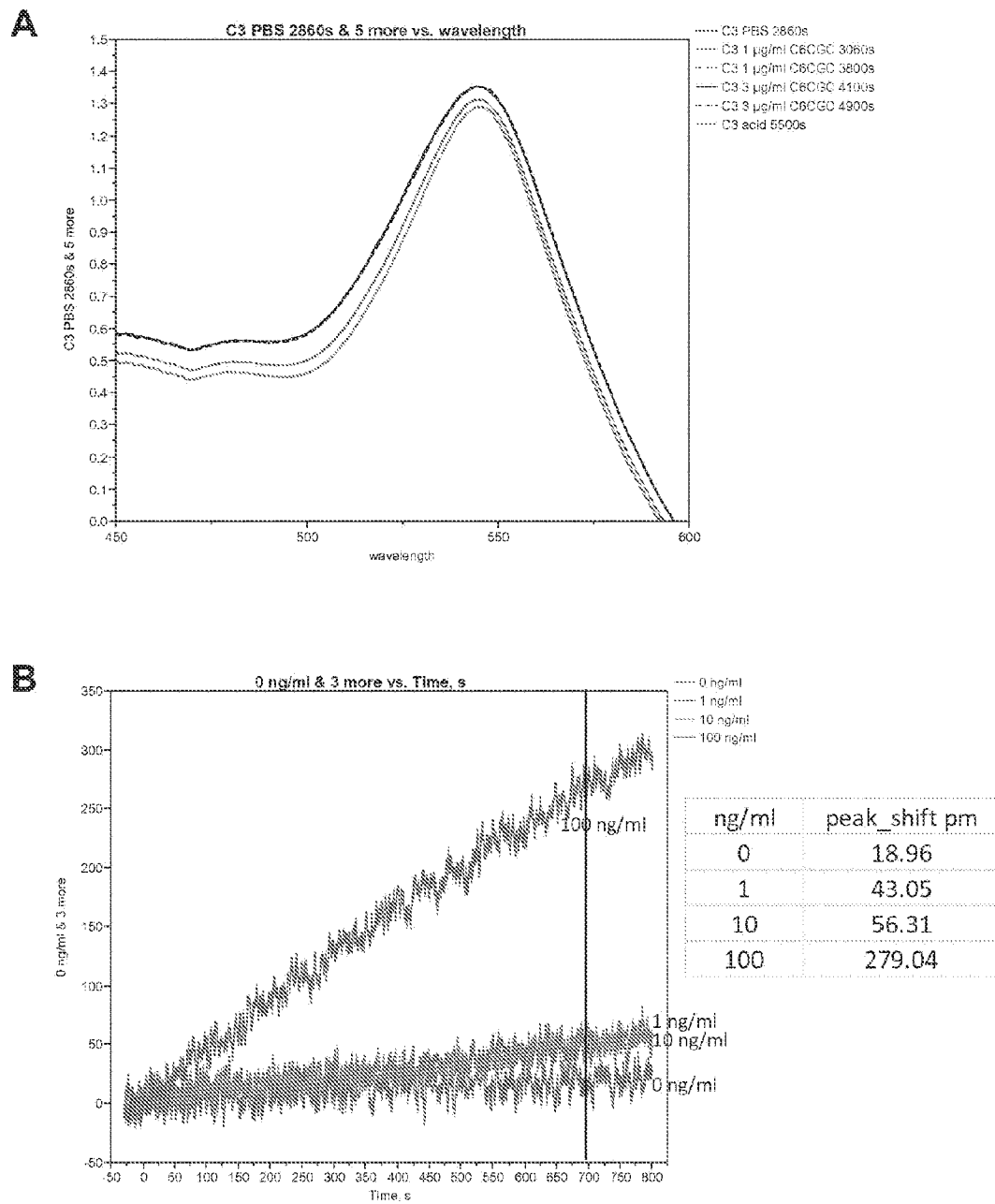
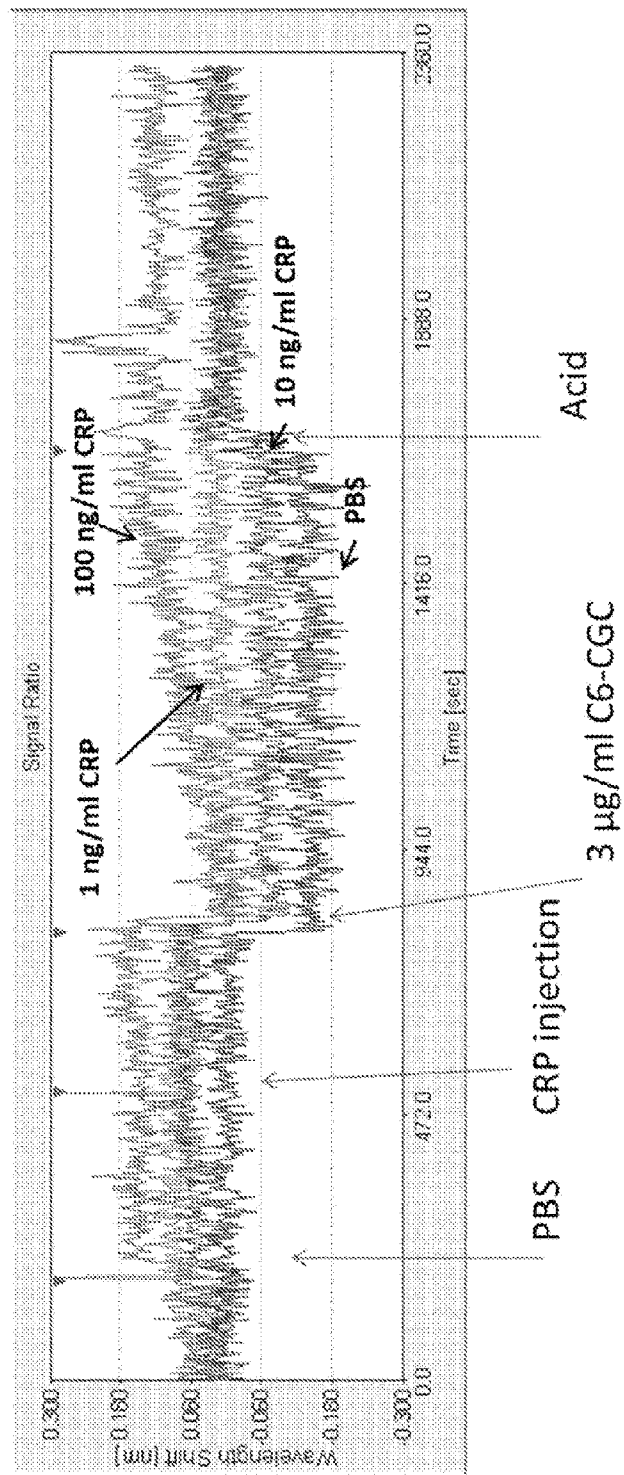
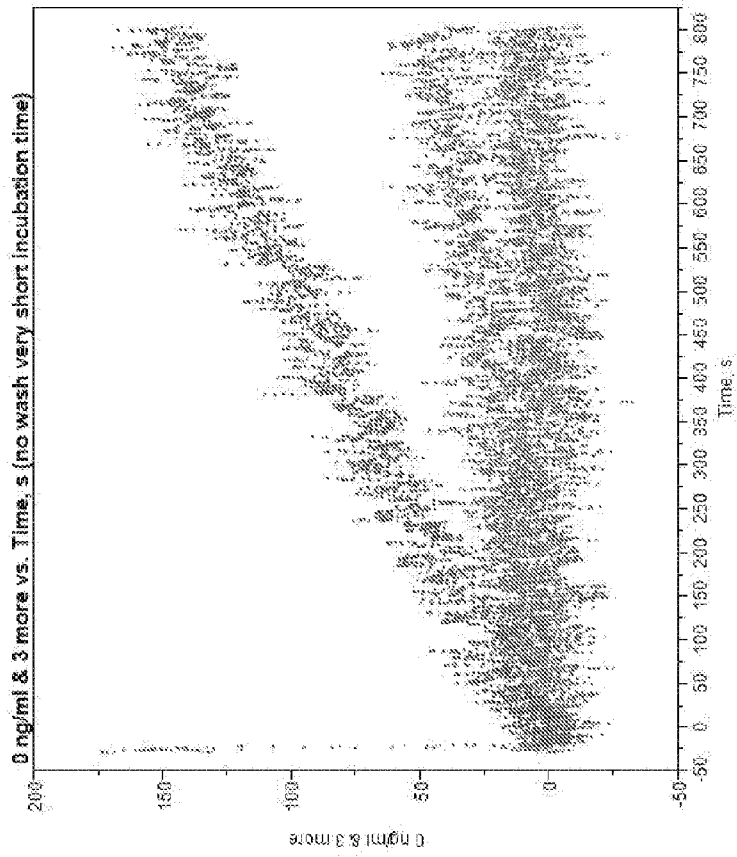


FIGURE 5



Blue= PBS; green= 1 ng/ml CRP; red= 10 ng/ml CRP and pink= 100 ng/ml CRP

FIGURE 6



× 0 ng/ml  
 × 10 ng/ml  
 × 10 ng/ml 2  
 × 100 ng/ml

100 ng/ml

Calibration Curve for Assay Time of 700 s			
15-20 min incubation		No wash/incubation	
ng/ml	peak shift pm	peak shift pm	peak shift pm
0	18.96		14.06
1	43.05		
10	56.31		43.38
100	279.04		137.44
			6.66

10 ng/ml

10 ng/ml 2  
0 ng/ml

FIGURE 7

