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(54) Title: ELECTROLYTE CONTAINING PHOSPHONIC ACID

(54) Bezeichnung: PHOSPHONSÄUREHALTIGER ELEKTROLYT

(57) Abstract: A method is disclosed for production of solutions of aminophosphonic acids and polymeric sulphonic acids in aprotic solvents. Membranes for membrane methodologies are produced from said solutions. Said membranes can also be doped with phosphoric acid.

(57) Zusammenfassung: Es wird ein Verfahren zur Herstellung von Lösungen aus basischen bzw. niedermolekularen Phosphonsäuren und sulfonierten Polymeren in aprotischen Lösungsmitteln beschrieben. Aus diesen Lösungen werden Membranen ftr Membranverfahren hergestellt. Diese Membranen können zusätzlich noch mit Phosphorsäure dotiert werden.

VERIFICATION OF TRANSLATION

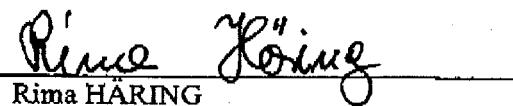
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I, Rima HÄRING, of FEIGENWEG 15, D-70619 STUTTGART, GERMANY

am the translator of the Patent Application No. PCT/DE2007/000133, which is the document of the corresponding AUSTRALIAN Patent Application and I state that the following is a true translation to the best of my knowledge and belief.

Signed this August 23, 2008.

Signature of translator:



Rima Häring

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Translation of WO 2007/031076

Titel:

~~Elektrolyte~~ containing phosphonic acid

Abstract:

A method is disclosed for production of solutions of aminophosphonic acids and polymeric sulfonic acids in aprotic solvents. Membranes for membrane methodologies are produced from said solutions. Said membranes can also be doped with phosphoric acid.

State-of-the-art:

Membranes from non-water-soluble sulfonated polymers that additionally contain low molecular phosphonic acids can not be manufactured by known post treatment methods of the sulfonated membrane.

The invention relates to a process to manufacture an electrolyte that solves this and other problems.

Description:

Low molecular phosphonic acids are not or only very poorly dissolved in organic solvents. This is especially the case for the aprotic solvents e.g. NMP, DMSO, DMF and DMAc.

Sulfonated polymers as sulfonated polysulphone or sulfonated polyetherketone are used as membranes in technology.

It would be desirable to blend these polymers with low molecular phosphonic acids. Thereby it is discovered that the phosphonic acids are not or only very poorly dissolved in organic solvents. This is especially the case for aminophosphonic acids. Experiments where membranes made of sulfonated polyetherketone were posttreated in the aqueous solution of these phosphonic acid revealed that the phosphonic acids penetrate the membrane only very marginally. This also does not change by increasing the temperature or the concentration of the aqueous phosphonic acid solution. Membranes are used whose ion exchange capacity is below the solubility in water. The limit of the solubility in water is depending on the used polymer approximately at 2.0 meq/gr. The limit of the solubility in water is at 2.4 meq/gr for polymers with a fluorinated or partially fluorinated backbone. Sulfonated PEEK does already start to dissolve at IEC of 1.8-1.85 in heated water. Hereinafter a process is presented that makes it possible to mix low molecular phosphonic acids with polymeric sulfonic acids in organic solvents and it is presented how particles, membranes, shapes and coatings can be manufactured.

This completely surprising and novel and new effect has been discovered.

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Aminotrimethylene-phosphonic acid is a low molecular aminophosphonic acid and not or only very marginally soluble in organic solvents especially in aprotic solvents as sulfolane, NMP, DMAc, DMF and DMSO. But this changes if a sulfonated polymer is dissolved in one of the aprotic solvents and the corresponding aminophosphonic acid is added. At the beginning the phosphonic acid is insoluble and dissolves slowly in the solution of the already dissolved polymer. The dissolution probably occurs because of an acid-base-interaction between the basic nitrogen and the sulfonic acid group. As the acidity of the sulfonic acid is higher than that of the phosphonic acid the aminophosphonic acid is dissolved in the acidic polymer. It is the immobilisation of a low molecular base that has an additional functional group, in this case the phosphonic acid. Every compound that is capable to form an ionic interaction with the sulfonic acid group of the polymer can be used as a base. The preferred solvent is DMSO. It was found out surprisingly that the calculated equivalent amounts of the aminophosphonic acids do not dissolve in a solution of sulfonated polyetherketone in NMP. This is especially the case for aminophosphonic acids that contain more than one NCP-group. One example therefore is diethylene-trisamino-pentamethylene-phosphonic acid (DTPMP). Theoretically to each molecule sulfonic acid on the polymer-chain maximum one molecule DTPMP can be added. But this does not function in NMP. In DMSO it is possible. The reason therefore is probably that NMP itself is a basic molecule. The amino group competes with the NMP (N-methylpyrrolidone) for the sulfonic acid groups.

One sulfonic acid group on the polymer chain can immobilise one low molecular basic molecule via the acid-base-interaction. Is this exceeded then the excess in the residual solvent is not soluble any more.

It does not make any difference in which order the components solvent, polymeric sulfonic acid and aminophosphonic acid are mixed. In any case the sulfonic acid dissolves first and then dissolves the normally insoluble aminophosphonic acid via acid-base-interaction.

The basic molecule is described by the general formula $R_2N-CR_2PO_3H_2$, whereby R is independently from another an alkyl-, aryl-, heteroaryl-, a carbon atom substituted at will or hydrogen. The substitution pattern has only the restriction that it does not cleave the acid-base-interaction between the nitrogen and the sulfonic acid. In the case of ATMP (aminotrimethylene-phosphonic acid) both moieties R at the nitrogen are identical and R is $-CH_2-PO_3H_2$ and R at the carbon is hydrogen. When the nitrogen is protonated, $R_2NH^+-CH_2-PO_3H_2$ is obtained and the acidity of the phosphonic acid moiety is strongly increased. This results in higher proton conductivity.

Preferred are also compounds with the general formula $R_2N-CR-(PO_3H_2)_2$ whereby R is again as defined above. One example for this class of compounds of bisphosphonic acids is dimethyleneamino-methylene-bisphosphonic acid (MAMDP) $(Me_2N)CH(PO_3H_2)_2$ and aminoethylene-bisphosphonic acid $(MeC(NH_2))(PO_3H_2)_2$, whereby Me is CH_3 .

In favour are also compounds whose group R_2N-CR- of the general formula $R_2N-CR_2PO_3H_2$ was closed to a heterocycle, e.g. pyridine, imidazole, triazole or which is part of a heterocyclic system. The order of the atoms NCP is determining the increase of acidity of the phosphonic acid.

These are examples for aminophosphonic acids which were used: aminotrimethylene-phosphonic acid, diethylenetriamino-pentamethylene-phosphonic acid, ethylenediamino-tetramethylene-phosphonic acid and pentaethylenehexamine-octakis-methylenephosphonic acid. This enumeration is not restricting and can be added at will by the person skilled in the art.

The aminophosphonic acids are dissolved in the solutions of the sulfonated polymers. All polymers that dissolve in the described aprotic solvents can be used as sulfonated polymer. The preferable IEC of the sulfonated polymers is between 0.5 and 5 milliequivalent per gram.

Example 1:

Sulfonated polyetherketone with an IEC of 1.8 milliequivalent is dissolved in DMSO. 50 gram of a 10% by weight solution is used. The solution contains 5 gram sulfonated polymer. Equivalent to the sulfonic acid group the aminophosphonic acid diethylenetriamino-pentamethylene-phosphonic acid (DTPMP) is added to the solution. Maximum almost 5 gram can be added. In the above described example 3 gram are added. Now the aminophosphonic acid dissolves.

The solution is casted as a thin film on a glass plate. The thickness of the film is approximately 600 μ . The solvent is evaporated at a temperature of approx. 100°C in the drying oven. This is an established process to manufacture membranes. After the solvent has been completely removed the film is carefully detached from the glass plate. This is easier when the plate is still warm or the foil is carefully humidified with a squirt bottle. It is then easier to detach it from the plate.

The obtained foil is then soaked in 80% phosphoric acid and left in the oven at 60°C. Thereby the foil incorporates phosphoric acid.

After the post treatment with phosphoric acid (PA) the membrane has proton conductivity as well at temperatures at about 20°C as also at higher temperatures.

Example 2:

Sulfonated polyetheretherketone with an IEC of 1.3 milliequivalent is dissolved in heated DMSO. 50 gram of a 10% by weight solution is used. The solution contains 5 gram sulfonated polymer. To this solution the aminophosphonic acid diethylene-triamino-pentamethylene-phosphonic acid (DTPMP) is added equivalently to the sulfonic acid group. Maximum almost 2 gram can be added, in this example 1.5 gram are added. Now the aminophosphonic acid dissolves.

Example 3:

Sulfonated polyetherketone (SPEK) with an IEC of 1.8 milliequivalent is dissolved in DMSO. A 15% by weight solution is used. To this solution the aminophosphonic acid diethylene-triamino-pentamethylene-phosphonic acid (DTPMP) is added equivalently to the sulfonic acid group. The solution is casted as a thin film on a glass plate. The thickness of the film is approximately 600 μ . The solvent is evaporated at a temperature of approx. 100°C in the drying oven. This is an established process to manufacture membranes. After the solvent has been completely removed the film is carefully detached from the glass plate. This is easier when the plate is still warm or the membrane is carefully humidified with a squirt bottle. Then it is easier to detach the foil from the plate.

The obtained foil is then soaked in 80% phosphoric acid and left in the oven at 60°C. Thereby the foil incorporates phosphoric acid.

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After the post treatment with phosphoric acid (PA) the membrane has a proton conductivity as well at temperatures at about 20°C as also at higher temperatures. The membrane is hygroscopic and absorbs eagerly water from the ambient air. The proton conductivity is based on the one side on the waterbased conduction mechanism and at anhydrous state it is based on the corporated phosphoric acid. When the membrane contains 1 gram SPEK and 1.03 gram DTPMP then the aminophosphonic acid is bound to the sulfonated polymer with exactly one nitrogen atom. Now the two remaining nitrogen atoms attach each a molecule phosphoric acid (PA). This results in the following ion exchange capacity: SPEK=1.8 and SPEK+DTPMP=9.75 and SPEK+DTPMP+PA=12.83. This means that the IEC of the starting polymer is strongly increased and thus also the proton conductivity.

Sulfonated polymers with immobilised amino phosphonic acids suit very well as electrolyte in the direct methanol fuel cell. The methanol cross over of such a membrane is clearly lower in comparison with the pure sulfonated polymers. The aminophosphonic acids that have more than one basic centre are hereby especially preferred. Examples for this is DTPMP and pentaethylenehexamine-octakis-methylene-phosphonic acid. If these phosphonic acids are used, a ratio of 1:2 with regard to the phosphonic acid molecule to the sulfonic acid groups is preferred. That means that one aminophosphonic acid forms an ionic bound to two sulfonic acids. Thus a cross linking results that improves the mechanical properties of the membrane. Hereby the IEC does not increase as strongly as at a ratio of 1:1 but still sufficient to obtain a clear increase of the proton conductivity. A cross linking is of course not possible if ATMP is used because this molecule has only one nitrogen atom.

1) Immobilisation of ATMP in a PBI membrane:

A foil of PBI (10 x 10 cm²) with a thickness of 60 μ is soaked in a 50% by weight solution of ATMP in water. The solution is left for 24 h at 80 °C in the oven. Then the membrane is removed and the surface is dried with pulp.

2) Immobilisation of PA in a PBI membrane:

A membrane of PBI (10 x 10 cm²) with a thickness of 60 μ is soaked in a 50% weight solution of PA in water. This solution is left for 24 h at 80 °C in the oven. The foil is dried as in example 1.

3) Immobilisation of PA and ATMP in a PBI membrane:

A membrane of PBI (10 x 10 cm²) with a thickness of 60 μ is soaked in an aqueous solution of ATMP and PA. The solution contains 25% weight ATMP and 25% weight PA. The solution is left for 24 h at 80 °C in the oven. The membrane is dried as in example 1.

4) Immobilisation of PA and ATMP in a PBI membrane:

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A membrane of PBI ($10 \times 10 \text{ cm}^2$) with a thickness of 60μ is soaked in an aqueous solution of ATMP and PA. The solution contains 25% weight ATMP and 25% weight PA. The solution is left for 24 h at 80°C in the oven. The membrane is dried at 130°C and then again soaked in the solution of ATMP and PA.

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Claims:

1. Process for preparing a solution of sulfonated polymer and a basic phosphonic acid characterised in that the polymer is dissolved in an aprotic solvent and that the basic phosphonic acid especially aminophosphonic acid is added to this solution.
2. Solution characterised in that it contains at least (A) any mixture of one or more low molecular phosphonic acids and (B) a sulfonated polymer, whereby the solvent is aprotic.
3. Use of the solutions according to claims 1 and 2 to produce membranes, particles, coatings and shapes.
4. Use of the membrane according to claim 3 in membrane processes especially in fuel cells.
5. PBI membrane characterised in that it contains any mixture of one or more low molecular phosphonic acids.
6. PBI membrane characterised in that it contains ATMP and phosphoric acid.
7. Process for doping of PBI characterised in that a PBI foil is soaked in subsequent steps or independently from another in a) a solution of diluted or concentrated phosphoric acid and/or b) a solution of ATMP in water or phosphoric acid.
8. PBI membrane containing immobilised ATMP, phosphoric acid and sulfonated polyetherketone.
9. Use of doped membranes according to claims 1 to 4 as membrane, especially in membrane fuel cells.

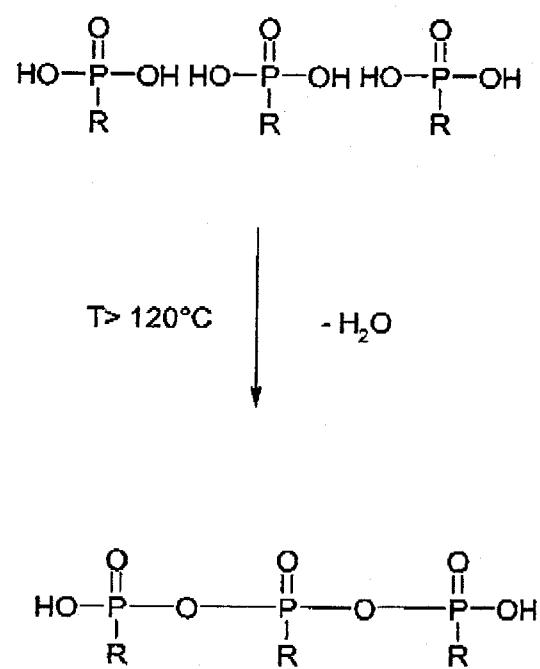
Figures:

Figure 1: condensation of phosphonic acids.

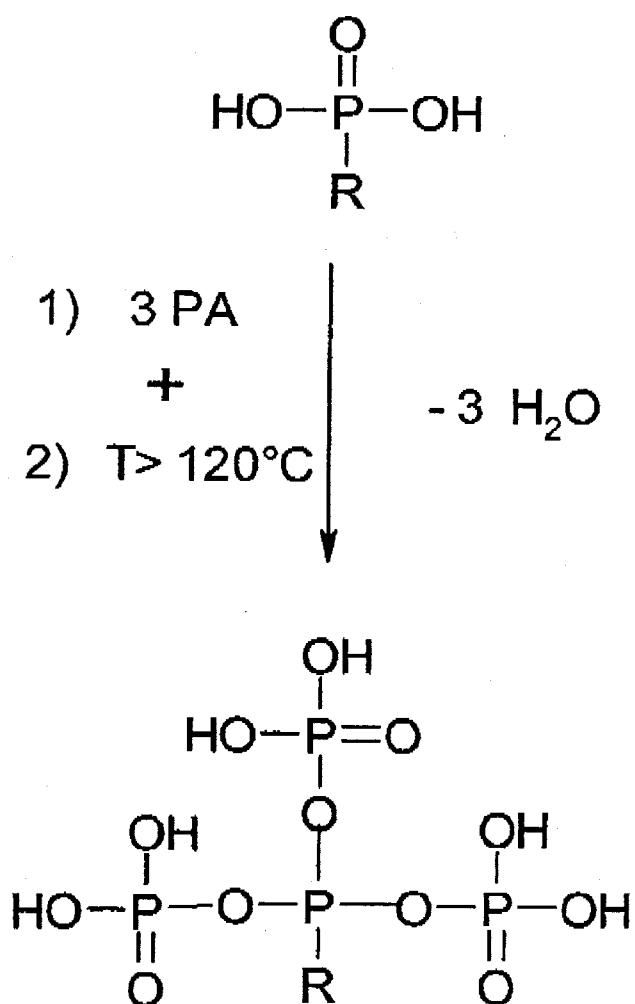


Figure 2: Mixed condensation of phosphonic acids with PA (general scheme)

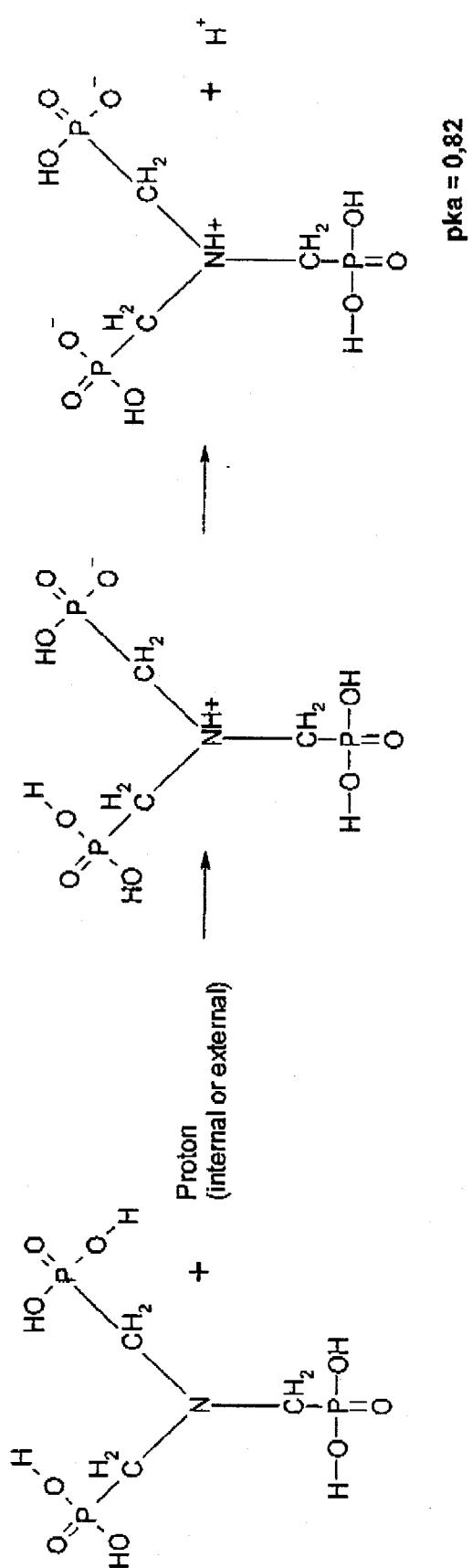


Figure 3: Interaction for increased acidity of ATMP here represented by internal protonation and the pK_a -value of deprotonation of the second phosphonic acid group.