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<p>(54) Title: A METHOD FOR ISOLATION AND PRODUCTION OF MAGNESIUM METAL, MAGNESIUM CHLORIDE, MAGNESITE AND MAGNESIUM BASED PRODUCTS</p>		
<p>(57) Abstract</p> <p>There is described a method for the isolation and production of magnesium metal, magnesium chloride and magnesium-based products from the leaching of laterite material, by isolating magnesium metal, magnesium chloride and magnesite from any step in the leaching of laterite material.</p>		

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A method for isolation and production of magnesium metal, magnesium chloride, magnesite and magnesium based products.

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The present invention relates to a method for the isolation and production of magnesium metal, magnesium chloride and magnesium-based products from the leaching of laterite material, the use of a method for the leaching of laterite material to isolate magnesium-based products, and magnesium-based products which are isolated according to the aforementioned method.

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Accordingly, a method is provided which produces magnesium and various magnesium-based products from the leaching of laterite materials.

15 JP 54155999 describes a method for introducing a gas containing CO₂ into a slurry of nickel-containing ore in order to extract magnesium oxide therefrom. The improvement comprises the addition of NaCl to the slurry to facilitate the extraction of magnesium oxide. The slurry is obtained by (1) calcination of nickel-containing ore at 500-800°C or selective reduction-calcination of nickel contained in the ore in order to obtain calcined ore, and the addition of calcined ore to water or (2) heating the ore to 20 500-800°C for selective reduction-calcination of the nickel contained therein in order to obtain calcined ore, contacting the calcined ore with an aqueous solution of ammonia containing free sulphur in order to extract useful metals such as nickel therefrom, and the addition of resulting extracted residue to the water. In one example, nickel ore was calcined at 500-700°C for one hour. The resulting calcined ore was added to an aqueous solution of 1 g/l NaCl in order to obtain a slurry. The waste gas containing 25 10% CO₂ was introduced into the slurry to extract magnesium oxide therefrom.

WO 81/02153 describes the production of magnesium oxide of high purity and high surface area. A method is described for the production of MgO from a solution of metal sulphate including MgSO₄ which comprises precipitating metals from the solution other than Mg, such as hydroxides, separating the solution from the precipitate, concentrating the separated solution to a specific gravity of from about 1.35 to 1.5 so that impurities including calcium sulphate are precipitated, separating the concentrated solution from the precipitate, isolating the dehydrated MgSO₄ crystals from the 35 concentrated solution and decomposing the dehydrated MgSO₄ crystals in order to form

MgO. Accordingly, a method is described for refining magnesium and nickel-containing ore comprising grinding the ore, preparing a solution of ground ore in sulphuric acid containing less than 10% water, adding water to the slurry of acid-ore in an amount effective to initiate a sulphating reaction, using the heat in the sulphating
5 reaction to heat the slurry of acid-ore, whereby water-soluble metal sulphates and insoluble residue are formed. Leaching the sulphating product with water to extract water-soluble metal sulphates in solution, followed by separation of the metal sulphate solution from the insoluble residue. Metals other than Mg are then precipitated as hydroxides. The solution is separated from the metal hydroxide precipitate, and the
10 solution is concentrated sufficiently to precipitate CaSO_4 . The concentrated solution is then separated from the precipitate and dehydrated MgSO_4 crystals are isolated from the concentrated solution. Dehydrated MgSO_4 crystals are broken down to form MgO having a controlled surface area.

15 DE 2906808 describes a method for isolating nickel from laterite ore, especially having a high magnesium content. This is done by leaching with sulphuric acid whereby the separated leaching solution is neutralised and the heavy metals are separated from the solution. The magnesium sulphate containing solution is at least partly evaporated and the crystallised magnesium sulphate separated. The separated magnesium sulphate is
20 heated in a reduced atmosphere during the formation of sulphur dioxide, water vapour and magnesium oxide, and sulphuric acid is produced from derived sulphur dioxide and is fed back into the leaching step.

US 5,571,308 describes a method for isolating nickel from laterite ore containing large
25 quantities of magnesium and iron. The ore is referred to as saprolitic ore and is subjected to leaching with a mineral acid from the group consisting of HCl, H_2SO_4 and HNO_3 . After leaching with HCl, the solution is separated from undissolved solid substances and nickel is preferably isolated by contacting the solution with a resin selective for nickel absorption. The residue contains iron and magnesium chloride
30 which may be subjected to pyrohydrolysis in order to produce their respective oxides and free HCl for recirculation into the leaching system. The nickel is extracted from the resin by using a stripper solution of said acid, and nickel is then extracted from the nickel-loaded stripper solution.

35 DE 3140380 describes a method for isolating nickel by a sulphuric acid leaching of a nickel-containing oxidic feedstock which also contains magnesium and iron, with the simultaneous production of magnesium oxide and cement. The method comprises

reacting the magnesium sulphate containing refine, at a pH of between 7 and 8.5, a temperature of less than 100°C and a magnesium sulphate concentration of less than 25%, with carbon dioxide and ammonia, wherein magnesium carbonate or hydroxide carbonate, as the case may be, is precipitated and an ammonium sulphate solution is formed. The precipitated product is separated from the ammonium sulphate solution and calcined to magnesium oxide, carbon dioxide being formed. Carbon dioxide is fed back to the precipitation step. A gypsum slurry is formed from the ammonium sulphate solution by adding burned calcium, and ammonia is released. Ammonia is then fed back to the precipitation step. Whilst additives are added, the gypsum slurry is dried and burned to give cement clinker. The sulphur-containing gas which is generated through the cement burning is converted in a sulphur factory to sulphuric acid and which at least partly is fed back into the feedstock leaching.

Consequently, the leaching of laterite material in order to isolate metals is not a new process. However, where nickel-containing laterites are concerned, the target metals from leached laterite materials have to date been nickel and cobalt. The leaching process used to isolate nickel and/or cobalt from nickel laterite ore also dissolves other metal ions. One of these ions is magnesium which is used *inter alia* in alloying with aluminium. With today's known technology, the magnesium production is a high-energy process.

Accordingly, the present invention relates to a method for the isolation and production of magnesium metal, magnesium chloride, magnesite and magnesium-based products from the leaching of laterite material, characterised in that magnesium metal, magnesium chloride and magnesite are isolated from any step during the leaching of the laterite material.

According to the method in the present invention, magnesium metal, magnesium chloride and magnesite are isolated after natural magnesium compounds in naturally occurring laterite have been dissolved.

The method according to the present invention is further characterised in that magnesite is produced in any step in the leaching of the laterite material by adding NaCO₃ and/or CaCO₃ to the leaching solution, after natural magnesium compounds in naturally occurring laterite have been dissolved. The use of a method for the leaching of laterite material to isolate magnesium-based products is also described.

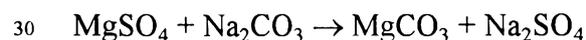
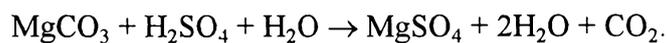
In known laterite leaching processes, magnesium oxide in the laterite material is kept at as low a level as possible because it is deemed to be a cost-increasing element in the processing and is also considered as waste. By using the present method, valuable magnesium-based by-products are produced from any step in the laterite process. The essential feature is to add sufficient acid to cause all the components in the material, including magnesium compounds, to dissolve so that desired ore minerals are then extracted from this solution. The magnesium-based products can be produced from any laterite leaching process. Consequently, the present invention makes it possible to isolate magnesium metal and other magnesium-based products which have previously gone to waste from known processes for leaching laterite material. Owing to the fact that MgO is considered to be an acid consumer in the leaching process and consequently leads to increased costs in the production of nickel and cobalt according to known methods for leaching laterite materials, it is at present kept at as low a level as possible. In the production of magnesium metal from laterite ore according to the present invention, it will be possible for the ore to contain larger amounts of MgO which in turn yield higher amounts of nickel.

Today, acid leaching technology is operated at different pressures and temperatures. In some cases atmospheric temperatures and pressures are used, whilst in other processes elevated pressures and temperatures are employed.

When magnesium is in solution from any of the various leaching processes, the production according to the present invention of identified magnesium products can be carried out.

25

When magnesium is in solution, according to the invention sodium carbonate (“soda ash”) is added and this produces magnesite



Below are exemplary embodiments showing examples of methods for isolating magnesium-based products from the leaching of laterite material according to the present invention.

35

EXAMPLE 1

A bulk sample of ore is mixed and dried in a furnace at 100°C. It is then mixed in a drum. From this, 1 kg of dry laterite ore is taken out and leached at atmospheric
5 pressure in H₂SO₄ at boiling temperature, 100°C, for eight hours. Water is added to 500 g of H₂SO₄ to obtain a total slurry volume of four litres. The leaching solution is filtered in a Buchner filter so that the liquid is filtered from the solid material. The acid liquid is then adjusted to a pH of 2-5-3.5 using NaOH and impurities are precipitated. Ni, Co and Mg are now in solution. The liquid is filtered again from the solid material.
10 500 ml of liquid is passed into a flask and stirred with a magnetic mixer. 25 g/l Na₂S is pipetted under the surface of the liquid. This results in the precipitation of Ni and Co from the solution with sodium sulphide as a source of hydrogen sulphide, and the formation of mixed sulphides of Ni and Co. The liquid is filtered once more from the solid material. The pH is increased to 4.5 by adding Na₂O₂. This precipitates out fewer
15 impurities and Mg is now in solution. The liquid is then filtered from the solid material. For the precipitation of MgCO₃, Na₂CO₃ is added and the pH is below 6.0. MgCO₃ is filtered and dried. For additional purification H₂SO₄ is added again and MgCO₃ is precipitated again with Na₂CO₃.

20 Production of MgO

MgCO₃ is calcined in a furnace at about 900°C. CO₂ is released and used again for precipitation of MgCO₃ with Na₂CO₃. This results in the precipitation of MgO.

Production of MgCl₂

25 MgO is dissolved in HCl in order to form a 35% MgCl₂ solution. Hydrated MgCl₂ is crystallised from the solution by using supersaturation on volume reduction due to boiling.

30 EXAMPLE 2

450 g dry laterite material containing 2.4% Mg is mixed with 225 g concentrated sulphuric acid and water is added in order to give a total volume of 1.5 litres. This slurry is leached under pressure for three hours at 255°C. Once the reaction is finished,
35 the pressure autoclave is cooled to room temperature and reacted slurry is removed. The slurry is then filtered in a Buchner filter to separate the solid material from the liquid. 500 ml of filtered liquid is placed in a glass container and stirred. To simulate

recirculation/neutralisation parameters, the liquid is doped with about 100 g of solid magnesium sulphate to produce an equivalent level consisting of 20 g/l Mg. The pH is increased to 2.0 by using an NaOH solution. Sodium sulphide solution is then added to remove Ni and Co as sulphides. The resulting slurry is filtered. The solution from the sulphide precipitation step is then treated with 5 g solid MgO to increase the pH to 4.5 and increase the Mg content of the solution. 30 g sodium peroxide is then added to the solution, whilst the pH is maintained at the same level by adding small quantities of sulphuric acid. This increases the ORP of the solution to about 655 mV in relation to AgCl and allows efficient precipitation of impurities such as iron. The resulting slurry is filtered.

The liquid from the preceding step is then treated with 670 ml of a solution consisting of 100 g/l sodium carbonate. This is added until a pH of 9.0 -9.5 is reached. Magnesium is effectively precipitated from the liquid as magnesium carbonate (magnesite). The resulting slurry is filtered to separate solid material from the liquid. About 30 g of dry magnesium carbonate was produced by using this procedure.

To produce magnesium chloride, 20 g magnesium carbonate is reacted with 75 ml of concentrated HCl. This mixture is heated and agitated in a container on a hot plate in order to dissolve the magnesium carbonate. The solution is then filtered. The clear filtrate is boiled in order to cause the magnesium chloride crystals to precipitate on account of supersaturation. The resulting slurry is filtered. 20 g solid magnesium chloride was produced by using this procedure.

To produce magnesium oxide, 5 g of magnesium carbonate is added to a scorification crucible. This material is then transferred to a furnace at 900°C for two hours. Carbon dioxide is removed and the remaining solid substance is magnesium oxide. As a result, about 2.40 g magnesium oxide was produced.

Consequently, the present invention describes a method for producing magnesium metal and magnesium-based products from the processing of laterite nickel ore. In the processing of laterite nickel ore, magnesium is in solution and precipitation according to the present invention can then result in the production of magnesium-based products and magnesium metal. This represents a considerable advance and involves great savings in costs in comparison with other methods for the production of magnesium. Magnesium metal is *inter alia* produced from native magnesite or sea water. The high costs involved in the production of magnesium metal are due to the fact that any

process requires the magnesium ion to be released from the host and often requires it to be dissolved. The fact that magnesium is dissolved in known processes for leaching laterite materials means that the magnesium production according to the present invention from waste substances is very favourable economically speaking. After
5 magnesium has been dissolved according to any conventional leaching process, it is precipitated by adding Na_2CO_3 and/or CaCO_3 or by electrodepositing or electrolysis, according to the present invention. Magnesite is used *inter alia* as an additive in cement, and magnesite brick is used in most industrial furnaces. In known leaching
10 processes, an attempt is made to avoid MgO in the feedstock because it increases the costs of production of nickel and cobalt. Changes in technology, chiefly in the autoclaving part of the process, have made the leaching process more favourable, economically speaking, for the production of nickel and cobalt, and consequently also for the production of magnesium metal, magnesium chloride, magnesite and
magnesium-based products according to the present invention.

P a t e n t c l a i m s

1.
A method for the isolation and production of magnesium metal, magnesium chloride,
5 magnesite and magnesium-based products from the leaching of laterite material,
characterised in that magnesium metal, magnesium chloride and magnesite are isolated
from any step in the leaching of laterite material.

2.
10 A method according to claim 1, characterised in that magnesium metal, magnesium
chloride and magnesite are isolated after natural magnesium compounds in naturally
occurring laterite have been dissolved.

3.
15 A method according to claim 1, characterised in that magnesite is produced in any step
in the leaching of laterite material by adding NaCO_3 and/or CaCO_3 to the leaching
solution after natural magnesium compounds in naturally occurring laterite have been
dissolved.

- 20 4.
A method according to claim 1, characterised in that magnesium chloride is produced
by the conversion of magnesite produced in claim 3 by adding acid, e.g., HCl, or
according to another chlorine reaction.

- 25 5.
A method according to claim 1, characterised in that magnesia is produced by calcining
magnesite produced according to claim 3 at temperatures which convert magnesite to
magnesia by converting MgCO_3 to MgO .

- 30 6.
A method according to claim 1, characterised in that magnesium metal is produced by
using magnesium-based products prepared according to claims 3, 4 and 5.

- 35 7.
Use of a method for leaching laterite material for isolating magnesium-based products.

8.

Magnesium-based products, characterised in that they are isolated according to the method in claims 1, 2, 3, 4 and 5.

AMENDED CLAIMS

[received by the International Bureau on 8 November 1999 (08, 11, 1999)
original claims 1-8 replaced by new claims 1-11 (2 pages)]

1.

A method for isolation and production of magnesite from leaching of laterite material,
5 characterised in that magnesite is formed by adding alkali and/or alkaline earth metal
carbonate to any leach stream in the leaching of laterite material after valuable metals
have been brought into solution.

2.

10 A method for isolation and production of magnesium chloride from leaching of laterite
material, characterised in that magnesium chloride is formed by concentrating any leach
stream in the leaching of laterite material with HCl after valuable metals have been
brought into solution.

15 3.

A method for isolation and production of magnesium chloride from leaching of laterite
material, characterised in that magnesium chloride is produced by conversion of the
magnesite produced in claim 1 by adding acid, e.g., HCl, or according to any other
chlorine reaction.

20

4.

A method for isolation and production of magnesia from leaching of laterite material,
characterised in that magnesia is produced by calcining the magnesite produced
according to claim 1 at temperatures which convert the magnesite to magnesia by
25 converting $MgCO_3$ to MgO .

5.

A method according to any of the claims 1-4, characterised in that the magnesium
chloride and the magnesite are isolated after natural magnesium compounds in naturally
30 occurring laterite have been dissolved.

6.

A method according to any of the claims 1-5, characterised in that the magnesium
chloride or the magnesite is isolated after nickel and cobalt has been removed from the
35 solution.

7.

A method according to any of the claims 1-5, characterised in that the magnesium chloride or the magnesite is isolated before nickel and cobalt are removed from the solution.

5

8.

Use of magnesium based products produced according to any of the claims 1-7 for the production of magnesium metal.

10

9.

Use of a method for leaching laterite material for isolating magnesium-based products.

10.

A method according to claim 1, characterised in that the carbonate is Na_2CO_3 .

15

11.

A method according to claim 1, characterised in that the carbonate is CaCO_3 .