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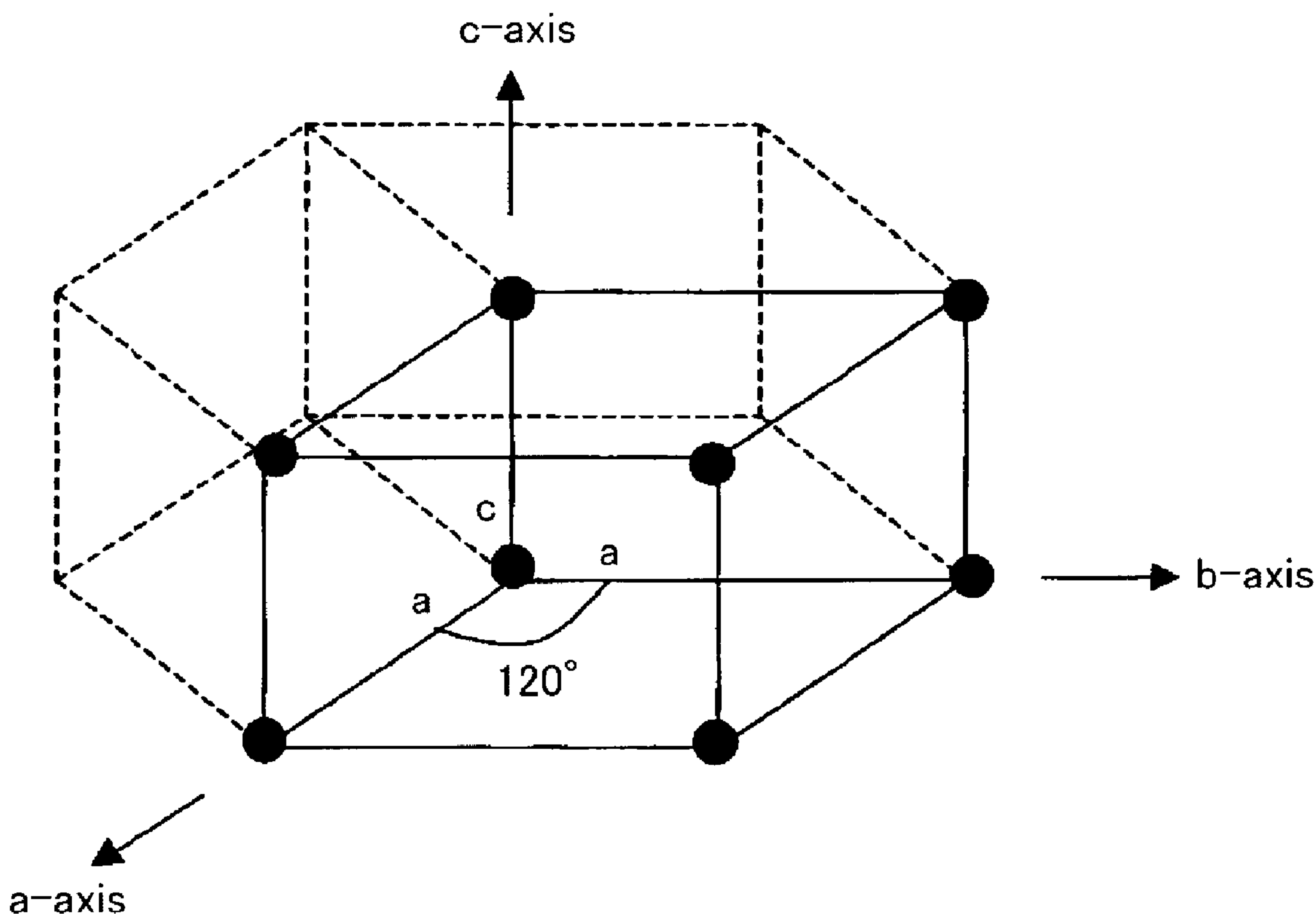
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(54) Title: PURIFICATION CATALYST FOR EXHAUST GAS, PRODUCTION METHOD THEREFOR, AND
PURIFICATION CATALYST EQUIPMENT FOR EXHAUST GAS



(57) Abrégé/Abstract:

A purification catalyst for exhaust gas which exhibits satisfactory performance even at a low temperature operation of starting or idling of engine (not more than 400°C), and a production method therefor are provided. The catalyst comprises an aluminum oxide supporting Pd, and the aluminum oxide is LnAlO_3 (Ln: rare-earth metal).

ABSTRACT

A purification catalyst for exhaust gas which exhibits satisfactory performance even at a low temperature operation of starting or idling of engine (not more than 400°C), and a production method therefor are provided.

The catalyst comprises an aluminum oxide supporting Pd, and the aluminum oxide is LnAlO_3 (Ln: rare-earth metal).

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DESCRIPTION

PURIFICATION CATALYST FOR EXHAUST GAS, PRODUCTION METHOD
THEREFOR, AND PURIFICATION CATALYST EQUIPMENT
FOR EXHAUST GAS

5 Technical Field

The present invention relates to a purification catalyst for exhaust gas, to a production method therefor, and to purification catalyst equipment for exhaust gas, and specifically relates to a production technique of a
10 purification catalyst for exhaust gas in which nitrogen oxides (NOx), hydrocarbons (HC), and carbon monoxide (CO) contained in an exhaust gas emitted from an internal combustion engine (for example, in a vehicle) can be simultaneously and effectively purified, thereby decreasing
15 undesirable components of the exhaust gas.

Background Art

For purifying exhaust gas containing, for example, CO, HC, and NO, precious metal elements (Pt, Rh, Pd, and Ir) are known to exhibit a high performance. Therefore, it is
20 preferable to employ the above-mentioned precious metal elements as the purification catalyst for exhaust gas. These precious metals are generally supported by Al_2O_3 which is a support having a high surface-to-weight ratio. On the other hand, composite oxides (for example, a perovskite-like
25 oxide) made by combining various elements have extremely varied properties. Therefore, it is preferable for a purification catalyst for exhaust gas to employ the above-mentioned composite oxides. Moreover, when the precious metal is supported by the composite oxides, the properties
30 of the precious metal are significantly changed. From this viewpoint, a preferable performance for purifying exhaust

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gas can be obtained in the purification catalyst for exhaust gas in which a precious metal is supported by a composite oxide.

Various catalyst mentioned above are now developed, and for example, a technique in which a coalescence rate of the precious metal can be reduced by employing a perovskite-like composite oxide as a support, judging from deterioration of the precious metal with a decrease of active sites by coagulation of the precious metal, is proposed (see the claims of Japanese Unexamined Patent Publication No. 5-86259). Moreover, another technique is proposed in which reduction of PdO can be suppressed by using a perovskite-like composite oxide in which the A site is defective, judging from reducing PdO which is an activated species in a NO reduction reaction, whereby the PdO changes to Pd whose activity is low, when the precious metal is Pd (see the claims of reacting of Japanese Unexamined Patent Publication No. 2003-175337).

Conventional purification catalysts for exhaust gas exhibit sufficient performance for removing CO, HC, and NOx contained in exhaust gas, while a vehicle is running, particularly while running at high temperatures (not less than 400°C). However, the conventional catalysts cannot exhibit sufficient performance for removing CO, HC, and NOx, in a vehicle at the starting or while idling at low temperatures (not more than 400°C).

As mentioned above, the reason why sufficient performance for purifying the exhaust gas cannot be obtained while running at low temperature is as follows. That is, in the conventional purification catalyst for exhaust gas, a precious metal, for example, Pt, Rh, or Pd, is supported on Al₂O₃ having a high surface-to-weight ratio, i.e., a high

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specific surface area. Due to the high surface-to-weight ratio of Al_2O_3 , the precious metal is advantageously supported in a highly dispersed condition. However, Al_2O_3 is a stable compound, and does not mutually affect the supported precious metal, whereby activity of the precious metal is not improved. Accordingly, sufficient performance during the running at low temperature cannot be obtained.

Moreover, while a vehicle is running, it is preferable for Pd to exist in the form of PdO which is highly reactive. However, even if Pd supported on Al_2O_3 initially exists in the form of PdO, it is reduced to a metal at high temperatures, whereby the activity is significantly decreased.

Disclosure of Invention

The invention was made in light of the above demands, and it is hence an object thereof to provide a purification catalyst for exhaust gas, in which activity of the precious metal is improved, and the decrease of activity at high temperatures is prevented, whereby sufficient performance even in a vehicle starting up or idling at low temperatures (not more than 400°C) can be obtained, and a production method therefor, and a purification catalyst equipment for exhaust gas.

The present inventors have intensively researched purification catalysts for exhaust gas, in which sufficient performance, even in a vehicle starting up or idling at low temperatures (not more than 400°C), can be exhibited. Consequently, it has been learned that a purification catalyst for exhaust gas made by supporting Pd on LnAlO_3 (Ln: rare-earth metal) has an effect of suppressing a reduction of PdO to Pd at a high temperature, whereby in the above-

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mentioned catalyst the high activity can be maintained during the running at low temperatures after running at high temperatures.

The present invention (the first aspect of the invention) was made in light of the above knowledge. That is, a purification catalyst for exhaust gas of the present invention is a catalyst in which Pd is supported on an aluminum oxide, and the oxide is LnAlO_3 (Ln: a rare-earth metal).

Moreover, the present inventors have also learned that LaAlO_3 among LnAlO_3 compounds, is trigonal or rhombohedral, and a B site in the perovskite-like composite oxide is Al in the LaAlO_3 , whereby a dipole moment of the LaAlO_3 is large, and an electric fluctuation of PdO bounded on the LaAlO_3 is larger than that of PdO which exists independently. Therefore, the oxidation state of Pd in a surface of the PdO supported is a state of Pd^{2+} over a large area. This state is a preferable state for purifying exhaust gas, whereby high activity at low temperatures can be obtained. Additionally, the present inventors have confirmed that this catalyst can exhibit high activity at low temperatures even after exposing the catalyst to operating conditions of about 1000°C .

The present invention (the second aspect of the invention) was made in light of the above knowledge. That is, in the above-mentioned purification catalyst for exhaust gas (the first aspect), it is preferable that the aluminum oxide be trigonal or rhombohedral.

Furthermore, the present inventors have also learned that when LnAlO_3 is produced, an aqueous nitrate solution of a component containing a carboxylic acid may be

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evaporated to dryness to obtain a carboxylic acid complex polymer, whereby LnAlO_3 is generated as a single phase, and a surface of the LnAlO_3 supporting Pd changes to a configuration in which an interaction with PdO is easy.

5 The present invention (the third and fourth aspects of the inventions) was made in light of the above knowledge. That is, in the above-mentioned purification catalysts for exhaust gas (the first and second aspects of the invention), it is preferable that at least one compound
10 selected from the group consisting of carboxylic acids having a hydroxyl group or a mercapto group and having a carbon number of 2 to 20, dicarboxylic acids having a carbon number of 2 or 3, and monocarboxylic acids having a carbon number of 1 to 20 is added to an aqueous nitrate solution
15 containing compounds of the elements of the catalyst, whereby a purification catalyst for exhaust gas is obtained (the third aspect of the invention). Moreover, in the purification catalysts for exhaust gas (the third aspect of the invention), it is preferable that the aqueous nitrate
20 solution is evaporated completely to obtain a carboxylic acid complex polymer, and that the carboxylic acid complex polymer is heated, whereby a

purification catalyst for exhaust gas is obtained (the fourth aspect of the invention).

As the carboxylic acid having a hydroxyl group or a mercapto group and having a carbon number of 2 to 20, oxycarboxylic acid and a compound in which an oxygen atom in the hydroxyl of the oxycarboxylic acid is replaced with a sulfur atom are cited. The carbon number of these carboxylic acids is 2 to 20 in light of solubility in water, is preferably 2 to 12, is more preferably 2 to 8, and is most preferably 2 to 6. Moreover, the carbon number of the monocarboxylic acid is 1 to 20 in light of solubility in water, is preferably 1 to 12, is more preferably 1 to 8, and is most preferably 1 to 6.

Furthermore, as concrete examples of the carboxylic acids having a hydroxyl group or a mercapto group and having a carbon number of 2 to 20, for example, glycolic acid, mercaptosuccinic acid, thioglycolic acid, lactic acid, β -hydroxy propionic acid, malic acid, tartaric acid, citric acid, isocitric acid, allo-citric acid, gluconic acid, glyoxylic acid, glyceric acid, mandelic acid, tropic acid, benzilic acid, and salicylic acid are cited. As concrete examples of the monocarboxylic acids, for example, formic acid, acetic acid, propionic acid, butyric acid, isobutyric acid, valeric acid, isovaleric acid, hexanoic acid, heptanoic acid, 2-methyl hexanoic acid, octanoic acid, 2-ethyl hexanoic acid, nonanoic acid, decanoic acid, and lauric acid are cited. In the above-mentioned acids, it is preferable to use acetic acid, oxalic acid, malonic acid, glycolic acid, lactic acid, malic acid, tartaric acid, glyoxylic acid, citric acid, gluconic acid, and more preferable to use oxalic acid,

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malonic acid, glycolic acid, lactic acid, malic acid, tartaric acid, glyoxylic acid, citric acid, gluconic acid.

Additionally, the present inventors have particularly researched in detail about a purification catalyst for exhaust gas in which rare-earth metal is applied to an A site of perovskite-like composite oxides. LnAlO_3 (Ln: a rare-earth metal) is trigonal or rhombohedral. Therefore, the electron state is extremely unstable. Moreover, in these oxides, Al is applied to a B site of perovskite-like composite oxides, whereby a dipole moment exists due to the strong covalent bond between Al and O. Therefore, the dipole moment of LnAlO_3 is larger than that of the conventional purification catalyst for exhaust gas, for example LaFeO_3 . Owing to the properties of LnAlO_3 , an electric fluctuation of PdO bounded on the LaAlO_3 is larger than that of PdO which exists independently, the oxidation state of Pd in a surface of the PdO which is supported is a Pd^{2+} state over a large area. Generally, Pd in the surface of the PdO exists in two states of Pd^{2+} and Pd^0 (metal state). In these states, a state of Pd^{2+} has higher activity as a purification catalyst for exhaust gas than a state of Pd^0 . That is, a purification catalyst for exhaust gas of the present invention, in which Pd is supported on the LnAlO_3 , has high activity, because most of the Pd on the surface of the PdO exists in a state of Pd^{2+} . Moreover, these catalysts can equally maintain the high activity state even after the catalysts are exposed during use to conditions of 1000°C .

The present invention (the fifth aspect of the invention) was made in light of the above knowledge. That is, in the above-mentioned purification catalysts for exhaust gas (the second to fourth aspects of the invention), it is preferable that Pd is supported on the LnAlO_3 (Ln: a

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rare-earth metal), and that Pd exist in a state of Pd^{2+} in the surface range that the Pd is supported (the fifth aspect of the invention).

Next, a production method for a purification
5 catalyst for exhaust gas of the present invention (the sixth aspect of the invention) is a method for preferably producing the above-mentioned catalysts (the first to fifth aspects of the invention). That is, the sixth aspect of the invention is a method in which when the purification
10 catalyst for exhaust gas in which Pd is supported on an aluminum oxide, at least one of compound selected from the group consisting of a (carboxylic acid having a hydroxyl group or a mercapto group and having a carbon number of 2 to 20, a dicarboxylic acid having a carbon number of 2 or 3,
15 and a monocarboxylic acid having a carbon number of 1 to 20 is added to an aqueous nitrate solution including a compound of the element of the catalyst, whereby a purification catalyst for exhaust gas is obtained.

In the above-mentioned production method for a
20 purification catalyst for exhaust gas (the sixth aspect of the invention), it is preferable that the aqueous nitrate solution be evaporated completely to obtain a carboxylic acid complex polymer, and that the carboxylic acid complex polymer be heated (the seventh aspect of the invention), and
25 it is more preferable that the heating temperature be not more than 1000°C (the eighth aspect of the invention).

Furthermore, purification catalyst equipment for exhaust gas (the ninth aspect of the invention), produced by using the above-mentioned purification catalysts for exhaust
30 gas (the first to fifth aspects of the invention), is desirable for internal combustion, for example, in a vehicle, particularly because nitrogen oxides (NO_x), hydrocarbon (HC),

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and carbon monoxide (CO) contained in an exhaust gas can be simultaneously and effectively removed or decreased by the equipment.

The purification catalyst for exhaust gas of the present invention in which Pd is supported on LnAlO_3 has a function in which the reduction of PdO to Pd metal can be suppressed. It is known that the shape of Ln (a rare-earth metal excluding Gd) variously changes in oxide states. For example, when a catalyst made by supporting Pd on La_2O_3 is exposed to high temperature conditions, La_2O_3 migrates onto the Pd grain from the contact area between Pd and La_2O_3 , whereby a shape is formed in which Pd particles are embedded in La_2O_3 , resulting in additional migration of minute amounts of La_2O_3 onto the Pd surface (Zhang et al., J. Phys. Chem., Vol. 100, No. 2, pp. 744-755, 1996). Even in the present system (LnAlO_3), Ln and Pd form a complex compound, whereby reduction of PdO to Pd metal can be suppressed. Owing to this effect, a purification catalyst for exhaust gas of the present invention can maintain the high activity state while running at low temperatures (not more than 400°C).

Moreover, in the LnAlO_3 , for example LaAlO_3 (including Pd/ PrAlO_3 or Pd/ NdAlO_3) is characterized in that the crystal system is trigonal or rhombohedral and the B site of the perovskite-like composite oxide is Al. The trigonal or rhombohedral system is, as shown in Fig. 1, a crystal system in which an ideal cubic system of a unit lattice is changed in the c-axis direction, and the angle between the a-axis and the b-axis is 120°. That is, the trigonal or rhombohedral system is a crystal system in which an ideal cubic system of a perovskite structure is significantly strained. In this crystal system, the electron state among constituent atoms is extremely unstable.

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Fig. 2 is a graph showing a XRD spectrum as data to confirm the differences of the crystal systems of the LaAlO_3 supporting Pd, etc. That is, when Pd/LaAlO_3 , Pd/PrAlO_3 , Pd/NdAlO_3 , and other perovskite-like composite oxides supporting Pd which are conventional purification catalysts for exhaust gas (Pd/GdAlO_3 , Pd/LaCoO_3 , Pd/LaFeO_3 , and Pd/LaMnO_3) are compared, a difference in strength of the main peak and deviance of position in other peaks are seen in the Fig. 2. Accordingly, judging from the fact that LaAlO_3 , PrAlO_3 , or AlO_3 is trigonal or rhombohedral, other perovskite-like composite oxides (GdAlO_3 , LaCoO_3 , LaFeO_3 , or LaMnO_3) are not trigonal or rhombohedral, but are rhombic. Additionally, among the conventional purification catalysts for exhaust gas, LaNiO_3 does not have a difference in strength at a main peak and deviance of position in other peaks from LaAlO_3 , whereby LaNiO_3 is trigonal or rhombohedral.

On the other hand, in the LaAlO_3 , PrAlO_3 , and NdAlO_3 , a B site in the perovskite-like composite oxide is Al, whereby the bond between Al and O has a high degree of a covalent bond. Therefore, some dipole moment is generated in a crystal of perovskite-like composite oxides which has generally a high degree of an ionic bond. As described above, the perovskite-like composite oxides, that is LaAlO_3 , PrAlO_3 , and NdAlO_3 , are trigonal or rhombohedral, and a B site in the perovskite-like composite oxides is Al in the oxides, whereby the dipole moment of the oxides is larger than that of the well-known purification catalyst for exhaust gas, for example LaFeO_3 .

Due to the dipole moment, an electric fluctuation of PdO bounded on the LaAlO_3 , PrAlO_3 , and NdAlO_3 is larger than that in which PdO exists independently. Therefore, the oxidation state of Pd in a surface of the PdO supported is a state of Pd^{2+} over a large area. There are two oxidation

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states of Pd in a surface of the PdO, which are a state of Pd²⁺ and a state of Pd⁰ (metal state). The state of Pd²⁺ has higher activity than the state of Pd⁰. That is, in the purification catalysts for exhaust gas of the present invention in which Pd is supported on the LaAlO₃, PrAlO₃, and NdAlO₃, the oxidation state of Pd in a surface of the PdO is the state of Pd²⁺, whereby the catalysts of the present invention have high activity. Moreover, the catalysts of the present invention can exhibit high activity during the running at low temperatures (not more than 400°C) even after exposing the catalyst to an operating condition of about 1000°C.

Furthermore, when the LaAlO₃, PrAlO₃, or NdAlO₃ is produced, an aqueous nitrate solution of compounds of the component elements containing a carboxylic acid is evaporated completely to obtain a carboxylic acid complex polymer, and the polymer is heated at a relatively low temperature of 800°C, whereby LaAlO₃, PrAlO₃, or NdAlO₃ is generated as a single phase. On the other hand, when the LaAlO₃, PrAlO₃, or NdAlO₃ is produced in other ways, for example, solid-phase reaction, LaAlO₃, PrAlO₃, or NdAlO₃ is not generated as a single phase even if the heating at a relatively high temperature of 1700°C is performed (see Rare Earth Science, Kagaku-Dojin Publishing Company, Inc, Ginya Adachi, p. 564). That is, LaAlO₃, PrAlO₃, or NdAlO₃ of the single phase can be synthesized at the above-mentioned low temperature by using a carboxylic acid. Therefore, a sufficient surface-to-weight ratio can be obtained, and the catalyst can be used in a state in which the surface of the crystal lattice is active. In the purification catalyst for exhaust gas made by supporting Pd on the LnAlO₃ by using the method of the present invention, a sufficient surface-to-weight ratio and a strong interaction between LnAlO₃ and Pd

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can be obtained, whereby high activity at low temperatures can be realized.

As mentioned above, the LnAlO_3 (Ln: rare-earth metal) is trigonal or rhombohedral, whereby the electron
5 state among constituent atoms in LnAlO_3 is extremely unstable, and the bond between Al and O is a strong covalent bond, whereby some dipole moment is generated. Therefore, most of Pd supported on these oxides exists in a state of Pd^{2+} . In order to confirm this, in the Pd/LaAlO_3 which is a
10 representative of the present invention and in the Pd/LaFeO_3 and $\text{Pd/Al}_2\text{O}_3$ which are representatives of the conventional catalysts, states of Pd in a PdO surface were examined by XPS. Generally, a peak position of the metal component (Pd^0) of Pd is 335.5 ± 0.3 eV. On the other hand, a peak position
15 of the ion component (Pd^{2+}) of Pd is 336.6 ± 0.4 eV. Judging from this fact and the results shown in Fig. 3, in the Pd/LaAlO_3 , there is a peak at a position which is equivalent to the Pd^{2+} , and in the Pd/LaFeO_3 and $\text{Pd/Al}_2\text{O}_3$, there is a peak at a position which is equivalent to the Pd^0 .
20 Accordingly, states of Pd in a

surface of the LaAlO_3 are mostly Pd^{2+} . Additionally, the states of Pd in a surface of the LaAlO_3 are confirmed as mentioned above, whereby states of Pd in a surface of the other LnAlO_3 (Ln: rare-earth metal), for example PrAlO_3 and NdAlO_3 are similarly estimated.

Brief Description of Drawings

Fig.1 is a perspective illustration showing a crystal system of LaAlO_3 constituting a purification catalyst for exhaust gas of the present invention.

Fig. 2 is a graph showing a XRD spectrum as data to confirm the differences of the crystal systems of the LaAlO_3 supporting Pd, etc.

Fig. 3 is a graph showing 3d orbital vicinity of Pd examined regarding the Pd states in a PdO surface by XPS, in the Pd/LaAlO_3 which is an example of the present invention and in the Pd/LaFeO_3 and $\text{Pd/Al}_2\text{O}_3$ which is an example of the conventional technique.

Best Mode for Carrying Out the Invention

Hereinafter, the present invention will be concretely explained by examples.

<Practical Examples 1 to 3>

[Production of composite oxides as support]

Predetermined amounts of lanthanum nitrate hexahydrate and aluminum nitrate nonahydrate were dissolved in ion-exchanged water, whereby a mixed solution was obtained. Next, a predetermined amount of malic acid was dissolved in

ion-exchanged water, whereby an aqueous malic acid solution was obtained. These two solutions were mixed, the obtained mixed solution was set on a hot plate with a stirrer, and the mixed solution was heated to 250°C and agitated by a stirring bar, whereby evaporation of water into vapor was performed, complete evaporation was performed, and the dried sample was crushed into powder by mortar and pestle. The crushed sample was moved to an aluminum crucible, the sample was heated to 350°C at a rate of 2.5°C/min in a muffle kiln, and a heat treatment was performed at 350°C for 3 hours. Owing to the heat treatment, a provisional heated substance in which malate and nitrate-nitrogen (nitrate salt and nitrate ion) were removed was obtained. After crushing the provisional heated substance into powder and mixing for 15 minutes by a mortar and pestle, the obtained mixture was set in the aluminum crucible again, the sample was heated to 800°C at a rate of 5°C/min in the muffle kiln, and a heat treatment was performed at 800°C for 10 hours. Owing to the heat treatment, a perovskite-like composite oxide of which the composition was LaAlO_3 was obtained. Moreover, perovskite-like composite oxides of which compositions are PrAlO_3 and NdAlO_3 were similarly obtained.

[Supporting of precious metal]

Next, a predetermined amount of palladium nitrate dehydrate was dissolved in ion-exchanged water, whereby an aqueous palladium nitrate solution was obtained. The palladium nitrate and a predetermined amount of LaAlO_3 , PrAlO_3 , or NdAlO_3 which was in powder form were set in a flask which was like an

eggplant, and the sample was completely dried in a hot water bath at 60°C while decreasing pressure in the flask by a rotary evaporator. After that, the sample was heated to 250°C at a rate of 2.5°C/min in a muffle kiln, was heated to 750°C at a rate of 5°C/min, and was held at 750°C for 3 hours. Due to these treatments, catalyst powders of the Practical Examples 1 to 3, of which the compositions were Pd/LaAlO₃, Pd/PrAlO₃, and Pd/NdAlO₃, in which PdO was impregnated and supported on the perovskite-like composite oxides, were obtained.

Surface-to-weight ratios for these catalyst powders are shown in Table 1.

Table 1

Sample No.	Composition	Surface-to-Weight Ratio (m ² /g)
Practical Example 1	Pd/LaAlO ₃	9
Practical Example 2	Pd/PrAlO ₃	8
Practical Example 3	Pd/NdAlO ₃	8
Comparative Example 1	Pd/Al ₂ O ₃	80
Comparative Example 2	Pd/GdAlO ₃	9
Comparative Example 3	Pd/LaNiO ₃	5
Comparative Example 4	Pd/LaMnO ₃	15
Comparative Example 5	Pd/LaCoO ₃	4
Comparative Example 6	Pd/LaFeO ₃	5
Comparative Example 7	Pd/LaAlO ₃	1

[Estimation of activity]

Next, initial activities and activities after endurance running were estimated for the obtained catalyst powders. The estimation was performed by flowing model exhaust gas of a vehicle into catalysts under conditions in which A/F (air-fuel ratio) was

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substantially 14.6 and SV (stroke volume) was 5000 h⁻¹. Endurance running was performed for 20 hours at an endurance running temperature of 900°C by using model exhaust gas in which A/F (air-fuel ratio) was substantially 14.6. These results are shown in Table 2 and Table 3. That is, Table 2 shows a temperature at which CO, HC, and NO are reduced by 50% in a temperature increase test of catalysts before the endurance running. Moreover, Table 3 shows a temperature at which CO, HC, and NO are reduced by 50% in a temperature increase test of catalysts after the endurance running.

Table 2

Sample No.	Composition	Temperature for 50% Reduction (°C)		
		CO	HC	NO
Practical Example 1	Pd/LaAlO ₃	254	260	197
Practical Example 2	Pd/PrAlO ₃	251	256	201
Practical Example 3	Pd/NdAlO ₃	258	264	201
Comparative Example 1	Pd/Al ₂ O ₃	288	295	302
Comparative Example 2	Pd/GdAlO ₃	273	280	213
Comparative Example 3	Pd/LaNiO ₃	299	315	217
Comparative Example 4	Pd/LaMnO ₃	281	299	204
Comparative Example 5	Pd/LaCoO ₃	305	320	233
Comparative Example 6	Pd/LaFeO ₃	300	305	241
Comparative Example 7	Pd/LaAlO ₃	292	301	235

Table 3

Sample No.	Composition	Endurance Running Temperature(°C)	Temperature for 50% Reduction (°C)		
			CO	HC	NO
Practical Example 1	Pd/LaAlO ₃	900	317	324	260
Practical Example 2	Pd/PrAlO ₃	900	312	323	263
Practical Example 3	Pd/NdAlO ₃	900	318	332	279
Comparative Example 1	Pd/Al ₂ O ₃	900	326	335	>400
Comparative Example 2	Pd/GdAlO ₃	800	339	355	>400
Comparative Example 3	Pd/LaNiO ₃	800	328	352	354
Comparative Example 4	Pd/LaMnO ₃	800	309	320	321
Comparative Example 5	Pd/LaCoO ₃	900	329	354	278
Comparative Example 6	Pd/LaFeO ₃	900	365	363	>400
Comparative Example 7	Pd/LaAlO ₃	900	332	361	295

<Comparative Example 1>

Pd/Al₂O₃ was produced in a manner similar to that of the Practical Example 1, and various estimations for activity were performed. The endurance running temperature was set at 900°C. The result was also shown in Table 1 to 3.

<Comparative Example 2>

Pd/GdAlO₃ was produced in a similar manner with the Practical Example 1. The crystal system of the GdAlO₃ is rhombic. Various estimates of activity were performed for this catalyst. The endurance running temperature was set at 900°C. The results are also shown in Tables 1 to 3.

<Comparative Example 3>

Pd/LaNiO₃ was produced in a manner similar to that of the Practical Example 1. The LaNiO₃ is trigonal or rhombohedral. Various estimations for activity were performed for this catalyst. The endurance running temperature was set at 800°C. The results are also shown in Tables 1 to 3.

<Comparative Example 4>

Pd/LaMnO₃ was produced in a manner similar to that of the Practical Example 1. The crystal system of the LaMnO₃ is rhombic. Various estimations for activity were performed for this catalyst. The endurance running temperature was set at 800°C. The results are also shown in Tables 1 to 3.

<Comparative Example 5>

Pd/LaCoO₃ was produced in a manner similar to that of the Practical Example 1. The crystal system of the LaCoO₃ is rhombic. Various estimations for activity were performed for this catalyst. The endurance running temperature was set at 800°C. The results are also shown in Tables 1 to 3.

<Comparative Example 6>

Pd/LaFeO₃ was produced in a manner similar to that of the Practical Example 1. The crystal system of the LaFeO₃ is rhombic. Various estimations for activity were performed for this catalyst. The endurance running temperature was set at 900°C. The results are also shown in Tables 1 to 3.

<Comparative Example 7>

A given amount of lanthanum oxide and aluminum oxide were mixed by mortar and pestle, the mixed sample was moved to an aluminum crucible, the sample was heated for 10 hours at 1100°C in a muffle kiln, and LaAlO₃ was obtained by solid-phase reaction. A precious metal was supported in a similar manner of the Practical Example 1 by using the LaAlO₃, whereby Pd/ LaAlO₃ was obtained. Various estimations for activity were performed for this catalyst. The endurance running temperature was set at 900°C. The results are also shown in Tables 1 to 3.

According to the Tables 2 and 3, the purification catalysts for exhaust gas of the Practical Example 1 to 3 exhibit excellent temperatures at which CO, HC, and NO are reduced by 50% at any time before and after the endurance running. The reason for this is that the purification catalysts for exhaust gas of the Practical Examples 1 to 3 are made by supporting Pd on the LaAlO_3 , PrAlO_3 , or NdAlO_3 , and these catalysts have a property of suppressing a reduction of PdO to Pd at high temperatures, whereby the high activity can be maintained in the running at low temperatures after a running at high temperatures in the catalysts. Moreover, the purification catalysts for exhaust gas of the Practical Examples 1 to 3 are trigonal or rhombohedral, and a B site in the perovskite-like composite oxide is Al in the catalysts of the Practical Examples 1 to 3, whereby dipole moment of the catalysts is large. Therefore, an electric fluctuation of PdO bounded on the LaAlO_3 , PrAlO_3 , or NdAlO_3 is larger than that of PdO which exists independently. Furthermore, in the purification catalysts for exhaust gas of the Practical Examples 1 to 3, LaAlO_3 , PrAlO_3 , or NdAlO_3 is produced, aqueous nitrate solution of element containing carboxylic acid is evaporated completely to obtain carboxylic acid complex polymer, whereby LaAlO_3 , PrAlO_3 , or NdAlO_3 is generated as a single phase, and a surface of the LaAlO_3 , PrAlO_3 , or NdAlO_3 supporting Pd take a form in which interaction with PdO is easy. Additionally, when the mixed solution is produced, malic acid can be used as mentioned above, and when citric acid and oxalic acid are similarly used, the same effect can be obtained.

On the other hand, the purification catalysts for exhaust gas

of the Comparative Examples 1 to 7 cannot exhibit an excellent temperature at which CO, HC, and NO are reduced by 50% at any time before and after the endurance running. The reason is as follows. That is, in the catalyst of the Comparative Example 1, Al_2O_3 is a stable compound, and Al_2O_3 does not mutually affect precious metal supported, whereby activity of Pd does not improve. In the catalyst of the Comparative Example 2, the crystal system is rhombic, whereby electrons state among constituent atoms is not more unstable compared with the case of trigonal or rhombohedral. In the catalyst of the Comparative Example 3, even though the crystal system is trigonal or rhombohedral, Al does not exist in a B site in the perovskite-like composite oxide, whereby it is difficult to generate some of dipole moment in a crystal of perovskite-like composite oxides which generally has a high degree of probability of being an ionic bond. In the catalyst of the Comparative Examples 4 to 6, the crystal systems are rhombic, whereby electrons states among constituent atoms are not more unstable compared with the case of trigonal or rhombohedral. In the catalyst of the Comparative Example 7, carboxylic acid is not used when the catalyst is produce, whereby LaAlO_3 cannot be generated as a single phase. Therefore, sufficient surface-to-weight ratio cannot be obtained, and the surface of the crystal lattice cannot be used in an active state.

The purification catalyst for exhaust gas of the present invention can be applied to an internal combustion engine of vehicles in which nitrogen oxide (NO_x), hydrocarbon (HC) and carbon monoxide (CO) in exhaust gas are required to be

simultaneously and effectively purified and reduced recently.

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CLAIMS:

1. A method for producing a purification catalyst for an exhaust gas, the catalyst being PdO and Pd supported on perovskite composite oxide LnAlO_3 in which Ln is a rare-earth metal, which method comprises:

(A) adding at least one carboxylic acid selected from the group consisting of (i) a carboxylic acid having a hydroxyl group or a mercapto group and having a carbon number of 2 to 20, (ii) a dicarboxylic acid having a carbon number of 2 to 3 and (iii) a monocarboxylic acid having a carbon number of 1 to 20, to an aqueous nitrate solution of compounds of Ln and Al to obtain a mixed solution;

(B) evaporating off water from the mixed solution to obtain a dried product;

(C) calcining the dried product at a temperature of not more than $1,000^\circ\text{C}$ to obtain the perovskite composite oxide LnAlO_3 ; and

(D) supporting Pd onto the perovskite composite oxide LnAlO_3 .

2. The method according to claim 1, wherein the temperature at which the dried product is heated in step (C) is about 800°C .

3. The method according to claim 1 or 2, wherein the carboxylic acid (i) is employed.

4. The method according to claim 3, wherein the carboxylic acid (i) is at least one member selected from the group consisting of glycolic acid, mercaptosuccinic acid, thioglycolic acid, lactic acid, β -hydroxypropionic acid, malic acid, tartaric acid, citric acid, isocitric acid,

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allo-citric acid, gluconic acid, glyoxylic acid, glyceric acid, mandelic acid, tropic acid, benzilic acid and salicylic acid.

5. The method according to claim 3, wherein the
5 carboxylic acid (i) is malic acid.

6. The method according to any one of claims 1 to 5,
wherein the evaporation in step (B) is at about 250°C to
obtain a dried sample; and the dried sample is crushed and
is then heated at about 350°C to obtain the dried product.

10 7. The method according to any one of claims 1 to 6,
wherein the rare-earth metal Ln is La.

8. The method according to any one of claims 1 to 6,
wherein the rare-earth metal Ln is Pr.

9. The method according to any one of claims 1 to 6,
15 wherein the rare-earth metal Ln is Nd.

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Fig. 1

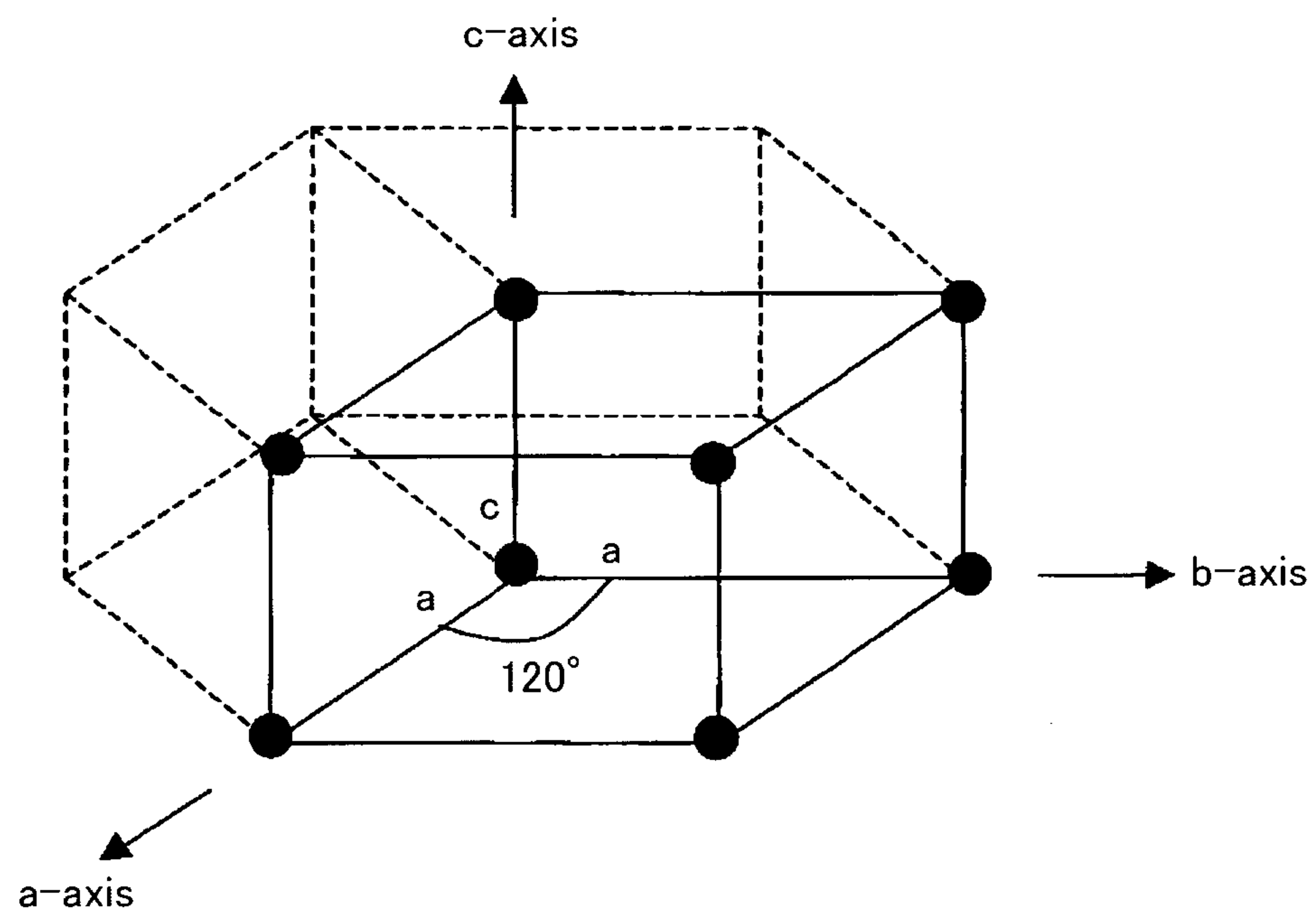


Fig. 2

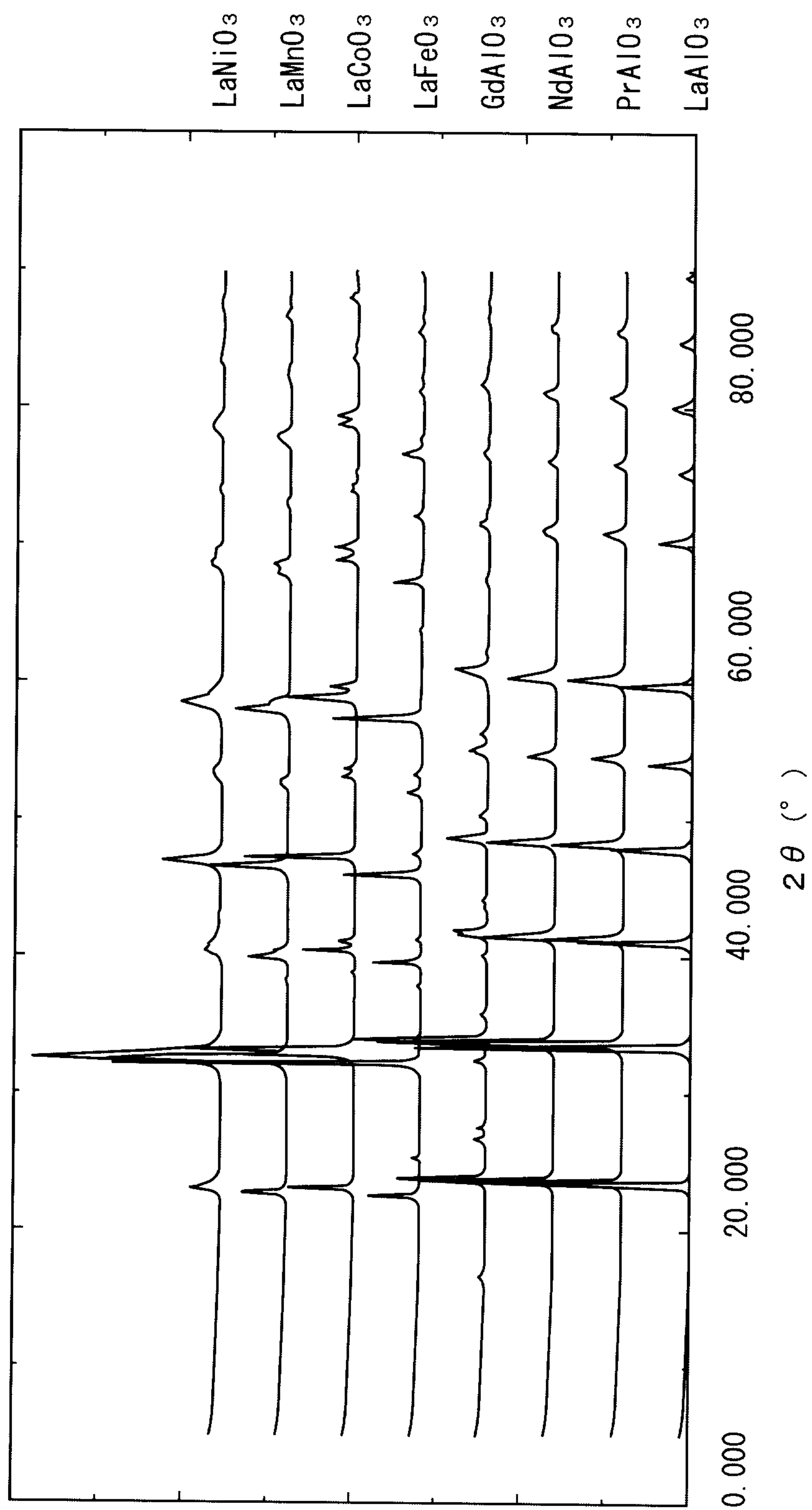
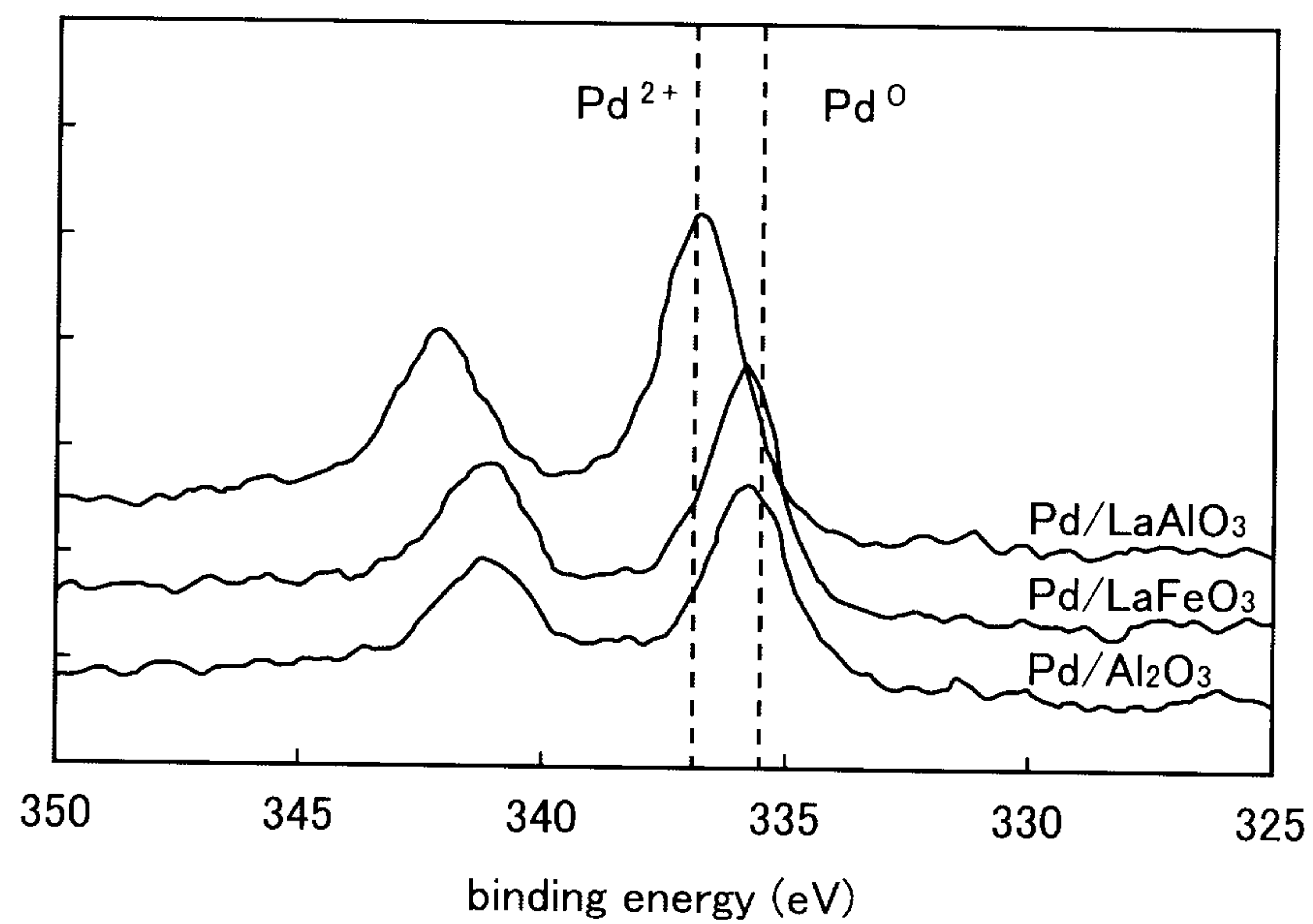
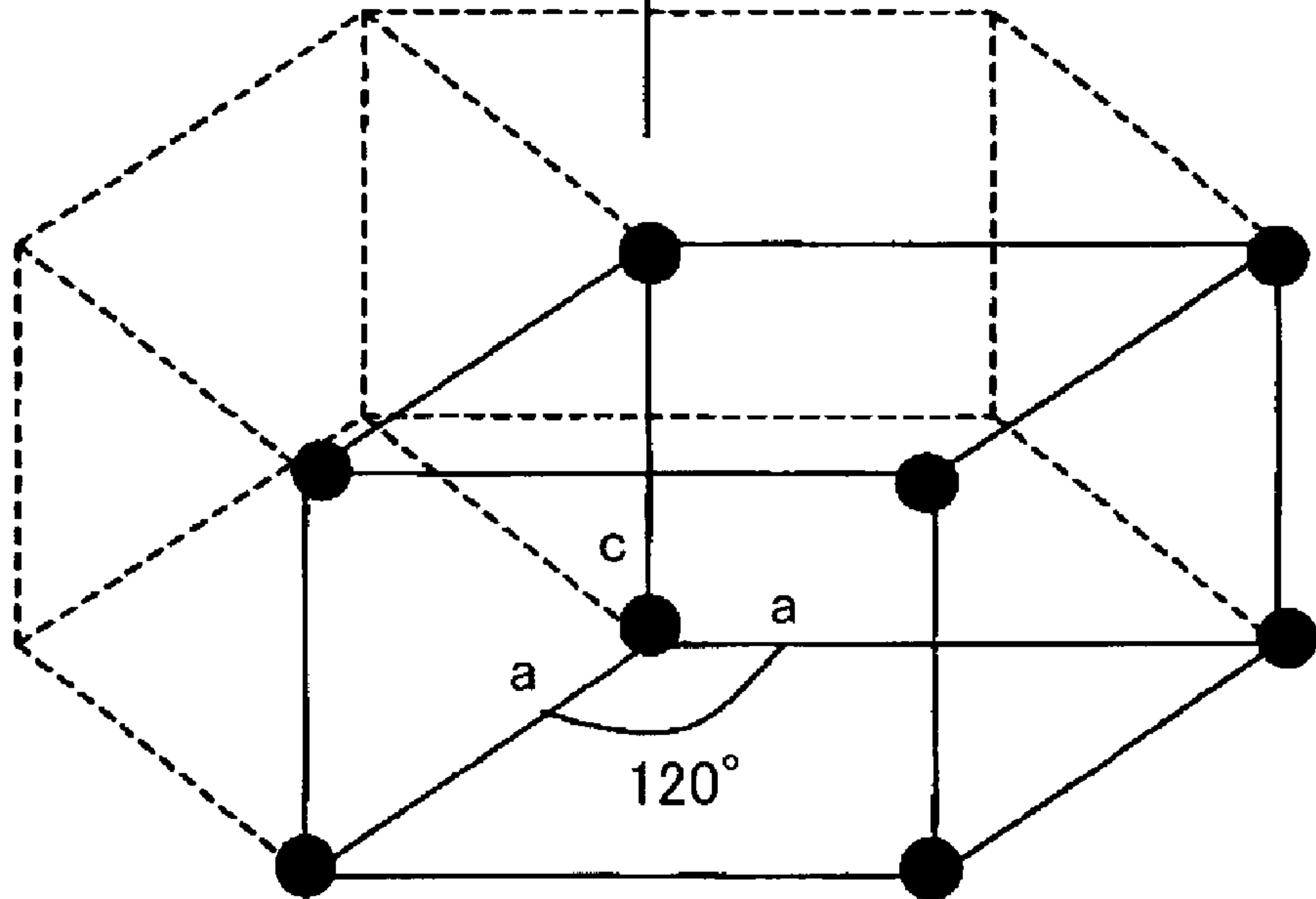


Fig. 3



c-axis



a

a

c

120°



b-axis



a-axis