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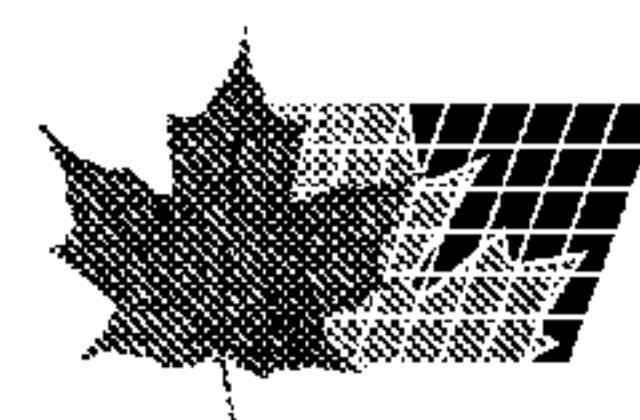
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(54) Title: PROCESS FOR THE PREPARATION OF TIOTROPIUM BROMIDE

(57) Abrégé/Abstract:

The invention is directed to improved processes for preparing Tiotropium bromide using scopoline salts.



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(54) Title: SCOPINE SALTS AND THEIR USE IN PROCESSES FOR THE PREPARATION OF N-DEMETHYL-TIOTROPIUM AND TIOTROPIUM BROMIDE

(57) Abstract: The invention is directed to improved processes for preparing Tiotropium bromide using scopine salts.

PROCESS FOR THE PREPARATION OF TIOTROPIUM BROMIDE**CROSS REFERENCE TO RELATED APPLICATIONS**

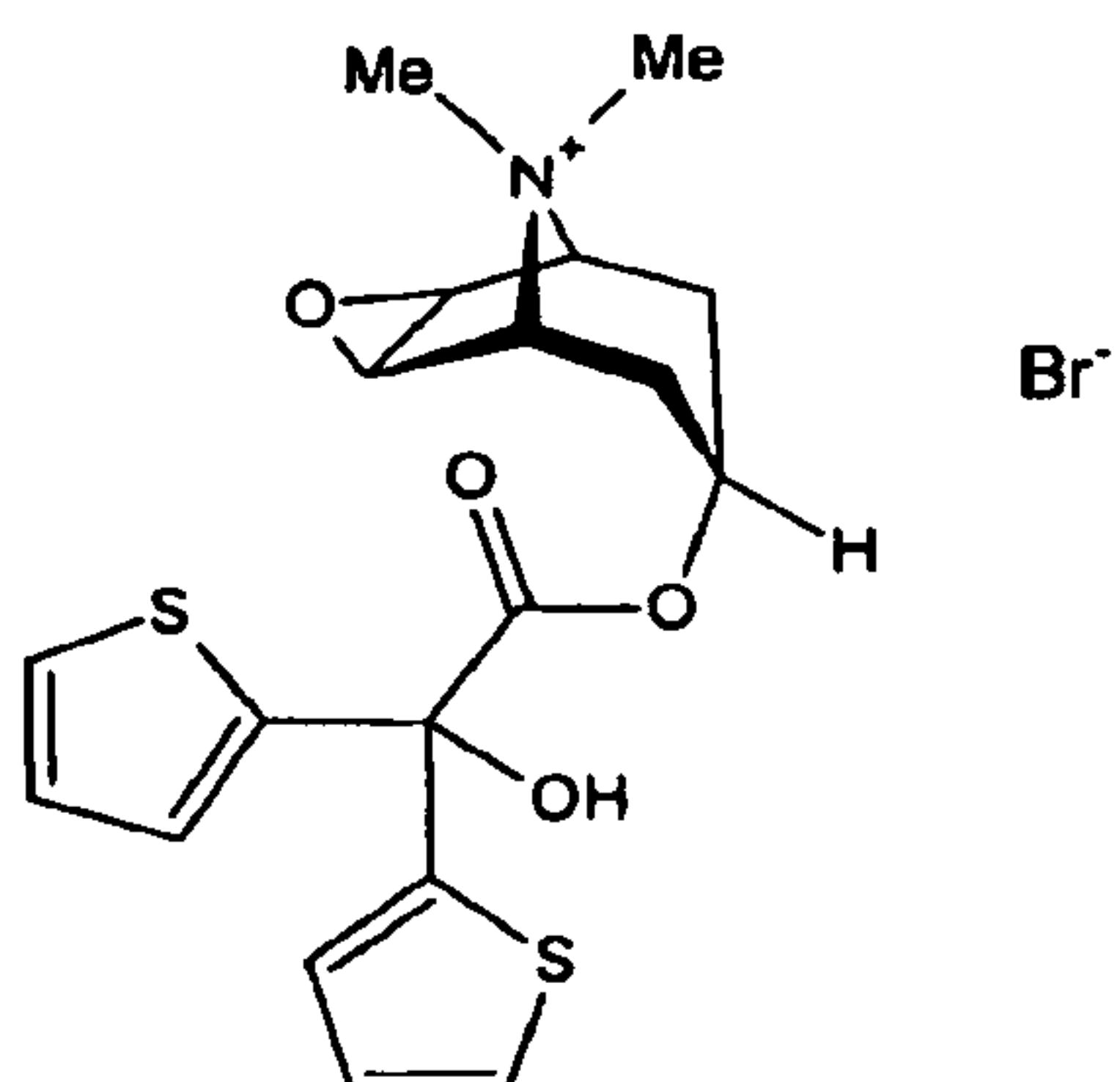
[0001] This application claims the benefit of the filing date of U.S. Provisional Patent Application No. 60/830,231 filed July 10, 2006; U.S. Provisional Patent Application No. 60/835,201, filed August 3, 2006; U.S. Provisional Application No. 60/835,200, filed August 3, 2006; and U.S. Provisional Application No. 60/836,037, filed August 7, 2006, the disclosures of which are hereby incorporated herein by reference. This application is also related to U.S. Patent Application No. 11/643,013, filed December 19, 2006.

FIELD OF THE INVENTION

[0002] The invention is directed to improved processes for preparing Tiotropium bromide.

BACKGROUND

[0003] Tiotropium bromide, $(1\alpha,2\beta,4\beta,5\alpha,7\beta)-7-[(\text{hydroxydi-2-thienylacetyl})\text{oxy}]-9,9\text{-dimethyl-3-oxa-9-azoniatricyclo[3.3.1.0]nonane bromide}$ or $6\beta,7\beta\text{-epoxy-3}\beta\text{-hydroxy-8-methyl-1}\alpha\text{H,5}\alpha\text{H-tropanium bromide, di-2-thienylglycolate}$, has the following chemical structure:



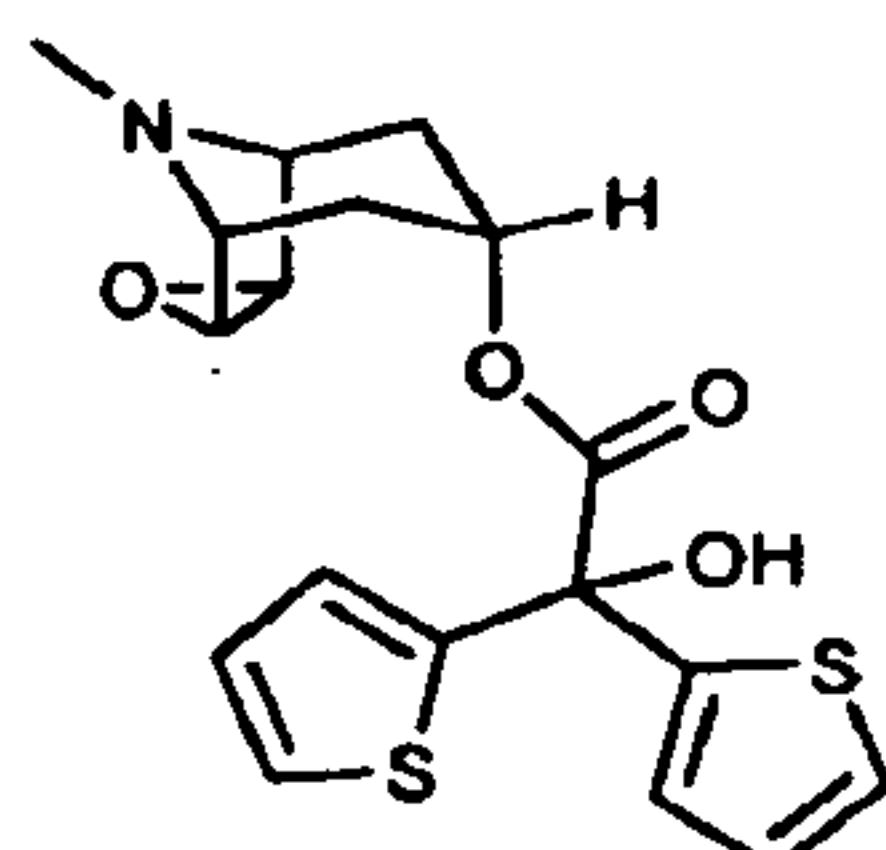
Tiotropium bromide



MW: 472.4

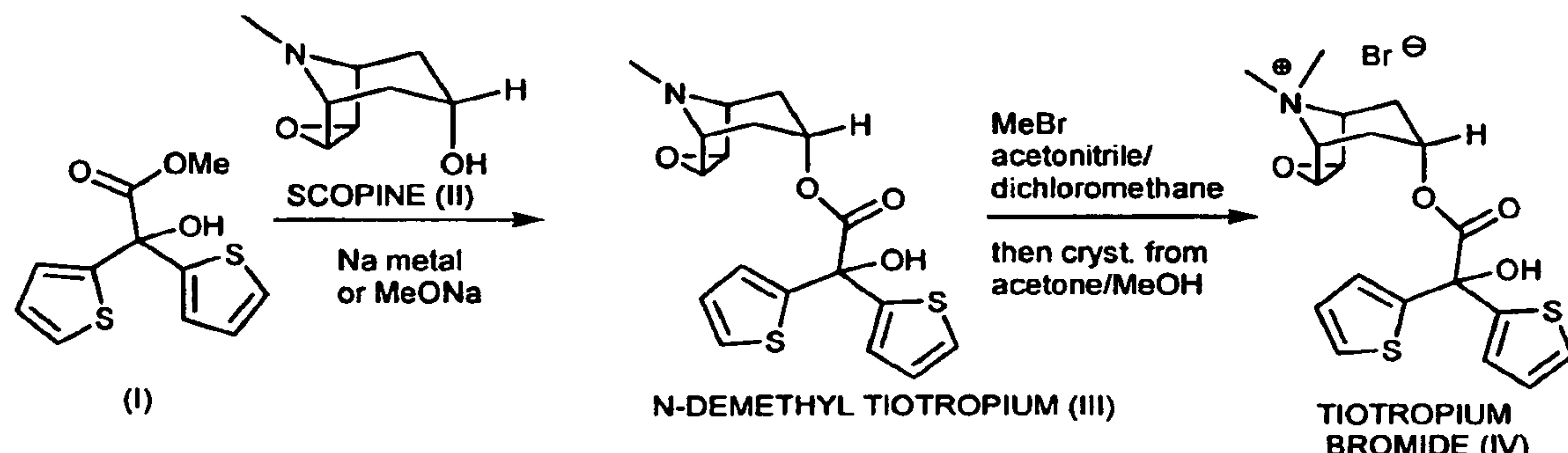
It is an anticholinergic drug with specificity for muscarinic receptors. As a bronchodilator it provides therapeutic benefit in the treatment of asthma or chronic obstructive pulmonary disease (COPD). This active pharmaceutical ingredient is administered by inhalation, and is available commercially as SPIRIVA® HandiHaler®.

[0004] Tiotropium bromide was first disclosed in United States Patent No. 5,610,163 where it was synthesized via N-demethyl tiotropium of formula III,



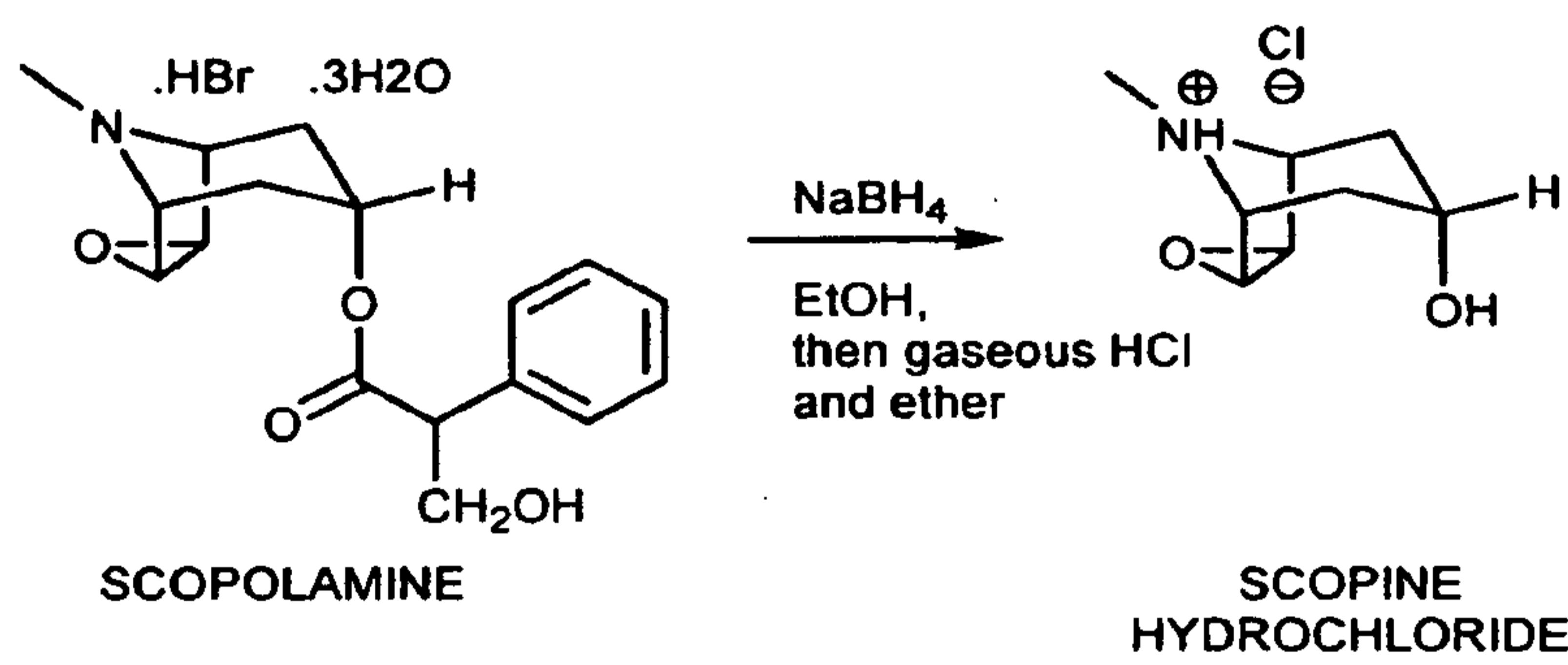
III

which was obtained by a reaction of methyl di(2-thienyl)glycolate of formula I and Scopine of formula II using sodium metal in melt or sodium methoxide in melt. Because of the dangerous reaction conditions, this method is not suitable for industrial scale preparation. The quaternization of N-demethyl-tiotropium is then carried out in a mixture of acetonitrile and methylene chloride using methyl bromide as a quaternizing agent. The process is illustrated in the following scheme:



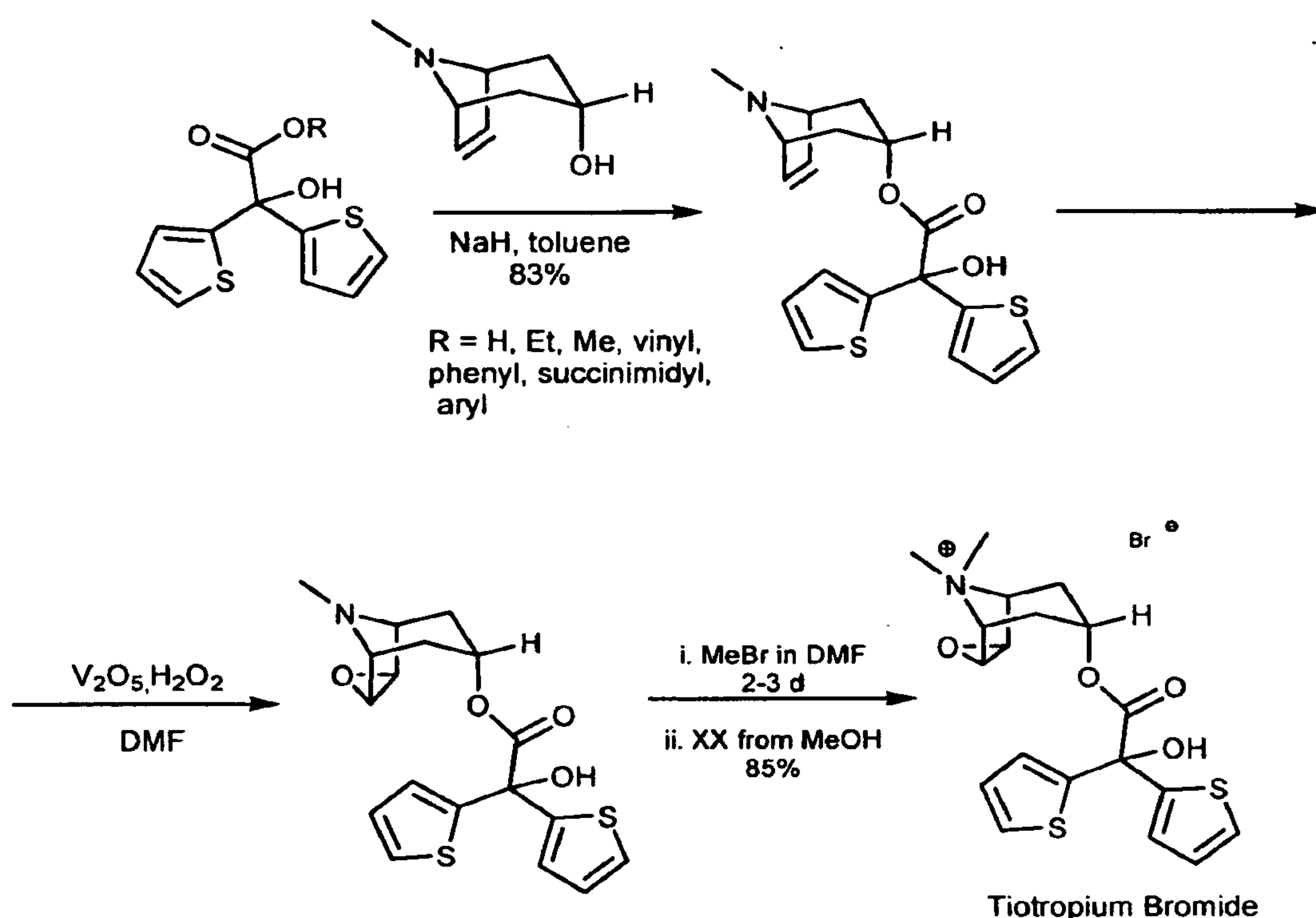
The product was then crystallized from a mixture of acetone and methanol.

[0005] Also described in the prior art is the preparation of Scopine HCl, which was first disclosed in GB 1469781, wherein it was prepared by reduction of scopolamine using sodium borohydride, followed by addition of HCl to the reaction mixture, a process illustrated by the following scheme:

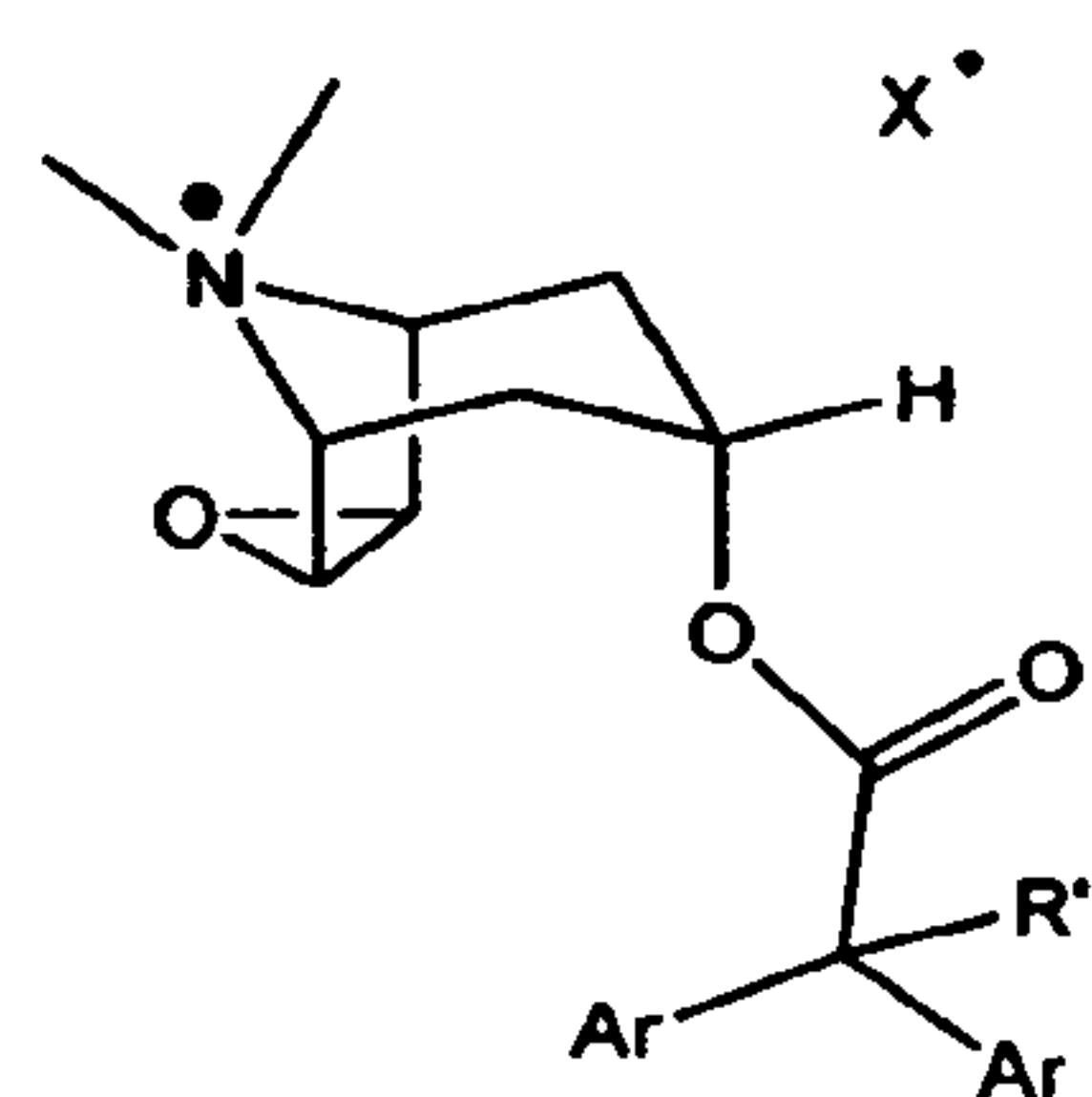


This salt of scopine can be used as a precursor for scopine base, however a process to remove inorganic salts from the desired product is not reported.

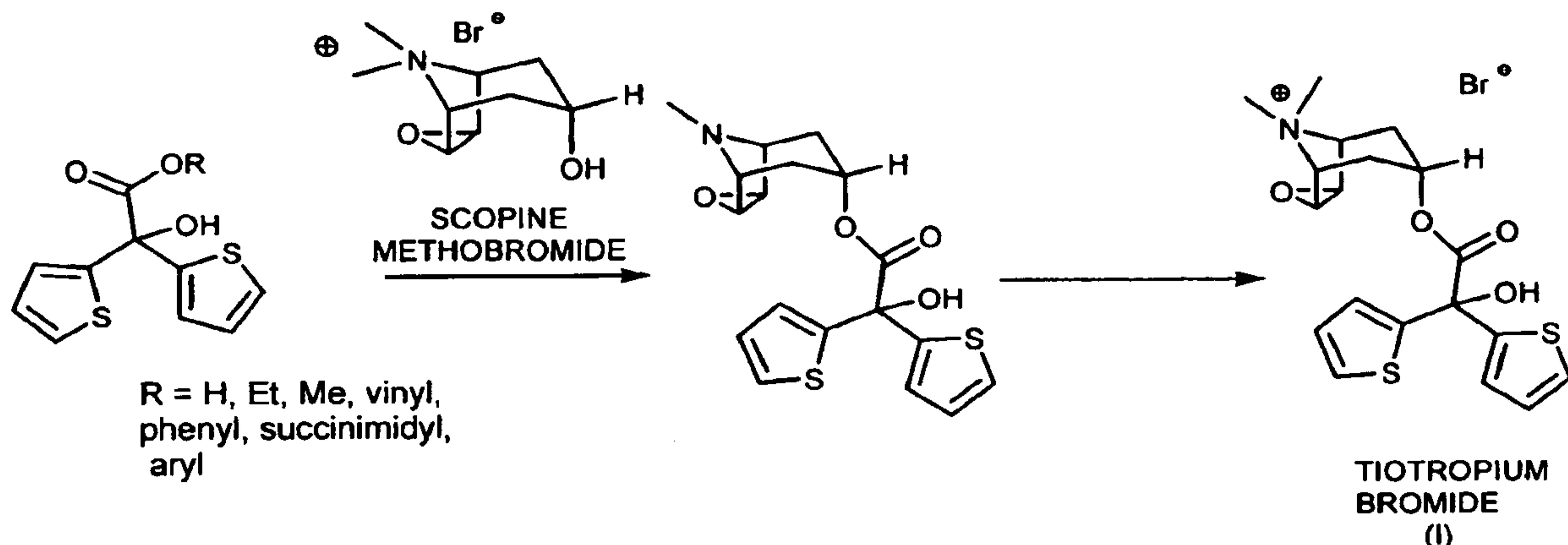
[0006] United States Patent Nos. 6,486,321 and 6,506,900 disclose a synthesis of Tiotropium and analogues via tropenol derivatives by introducing an additional epoxidation step, as described by the following scheme.



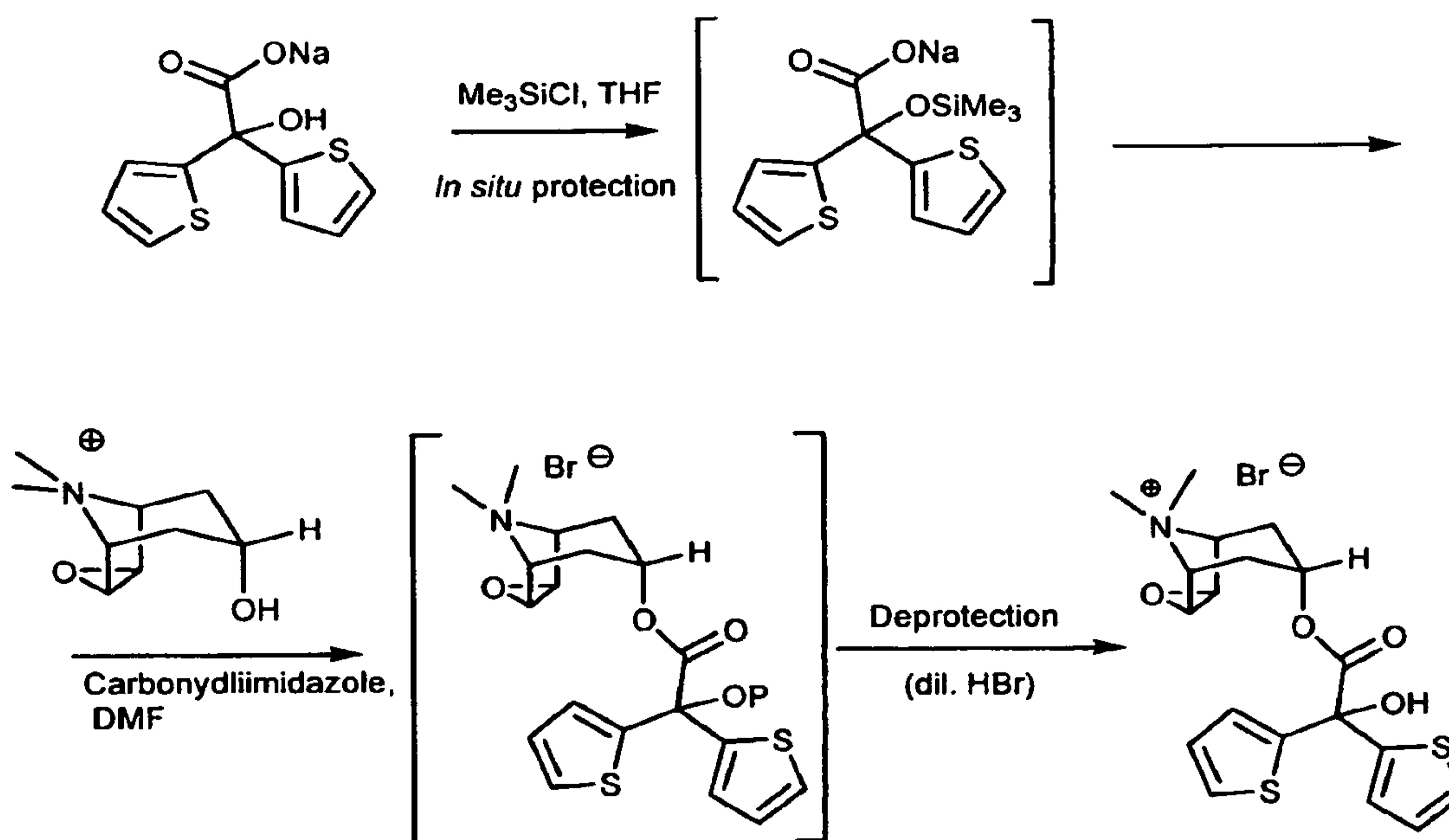
[0007] United States Patent No. 6,747,154, refers to formal approaches by stating "These processes known in the art may also be used to prepare the compounds of formula 1.



[0008] However, these methods of synthesis are more complex procedures involving a number of synthetic steps." Therefore a different synthetic approach was developed, where the coupling is carried out using scopoline methobromide rather than scopoline, but details, including the yield, of this coupling reaction are not reported.



[0009] United States Patent Publication No 2006/0047120 describes yet another approach, coupling scopine methobromide with trimethylsilyl-protected sodium dithienyl glycolate which is obtained in situ.

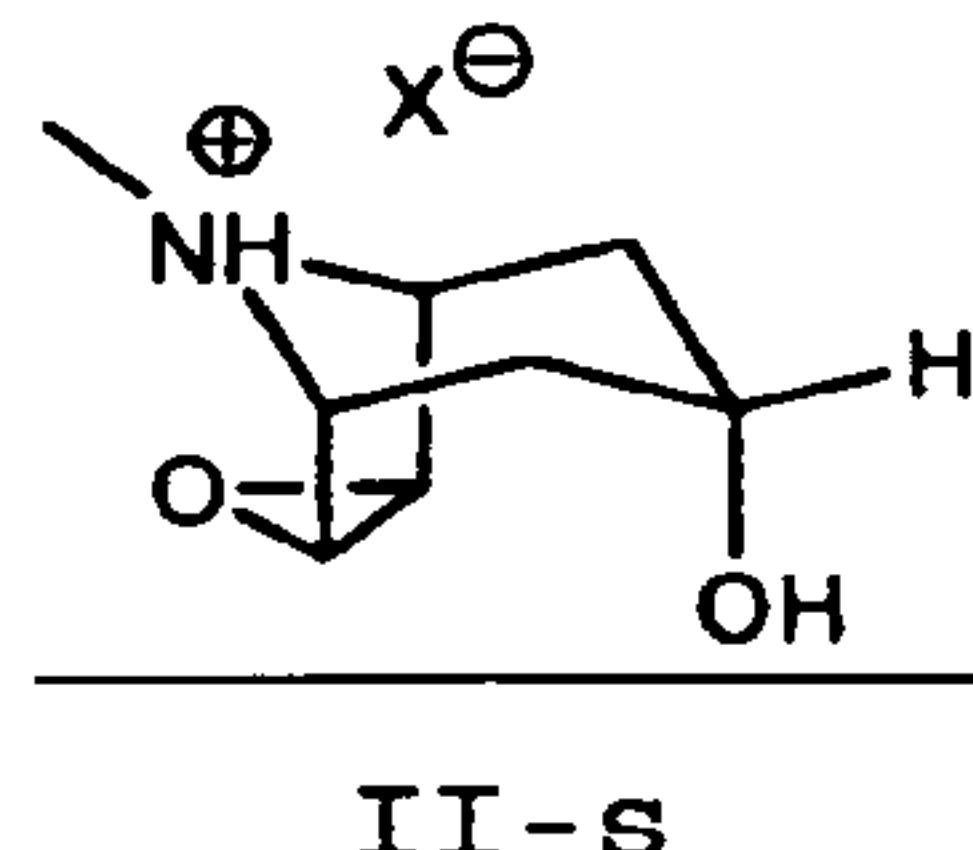


[0010] This application, provides that this approach was developed to improve the prior art synthesis for Tiotropium bromide.

[0011] Hence, an improved process to prepare Tiotropium bromide is needed.

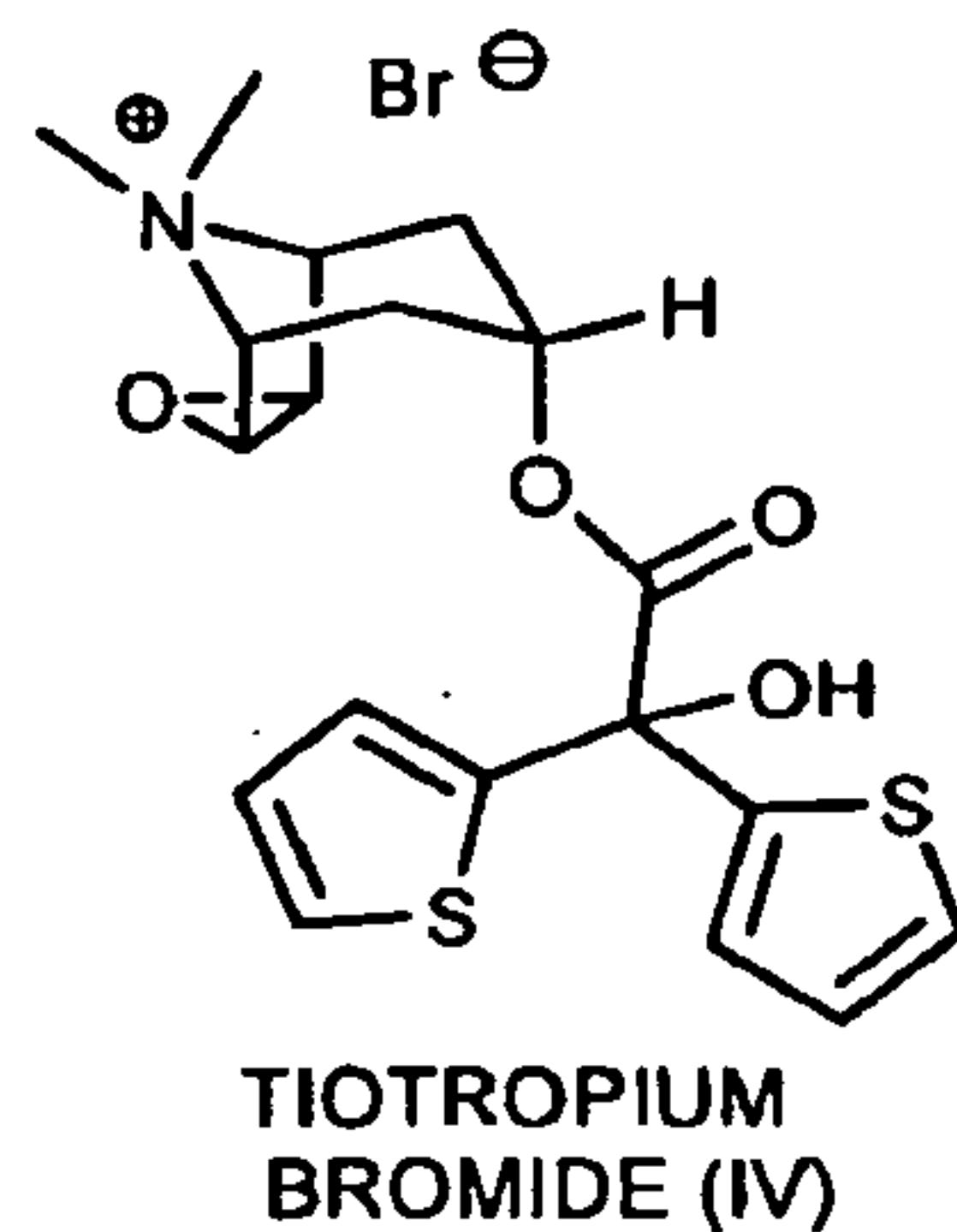
SUMMARY OF THE INVENTION

[0012] In one embodiment, the present invention encompasses scopine salt of formula II-s :

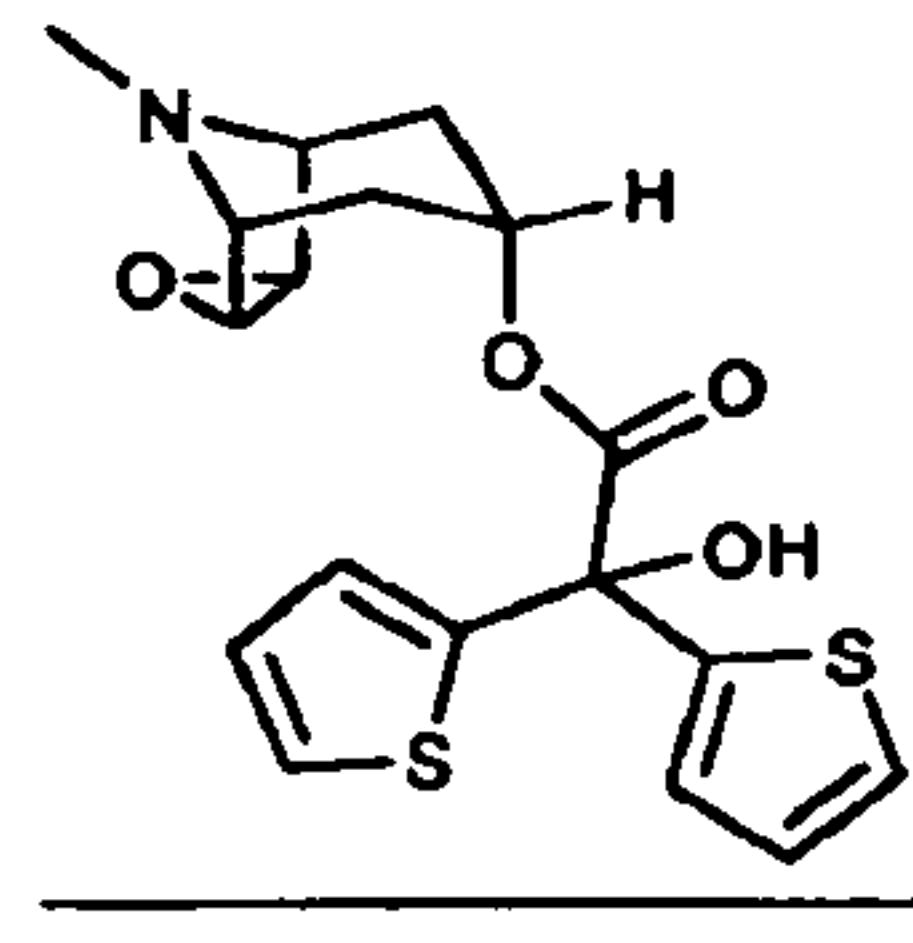


containing about 0.5% to about 40% by weight of salts; wherein X is Br, Cl, SO₄, MeCOO, PO₄, MeSO₃, tartrate, fumarate, citrate, maleate, succinate, p-toluene sulphonate or amidosulphonate.

[0013] In another embodiment, the present invention encompasses a process to convert scopine salt of formula II-s containing about 0.5% to about 40% by weight of salts to Tiotropium bromide of formula IV.



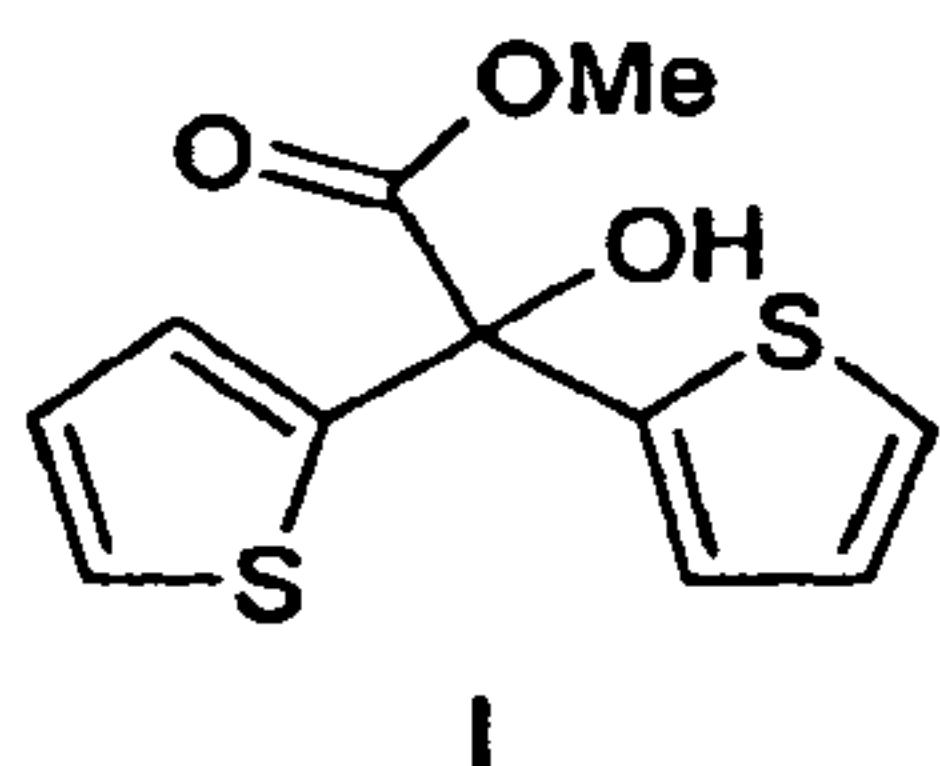
[0014] In yet another embodiment, the present invention encompasses the preparation of N-demethyl-tiotropium of formula III,



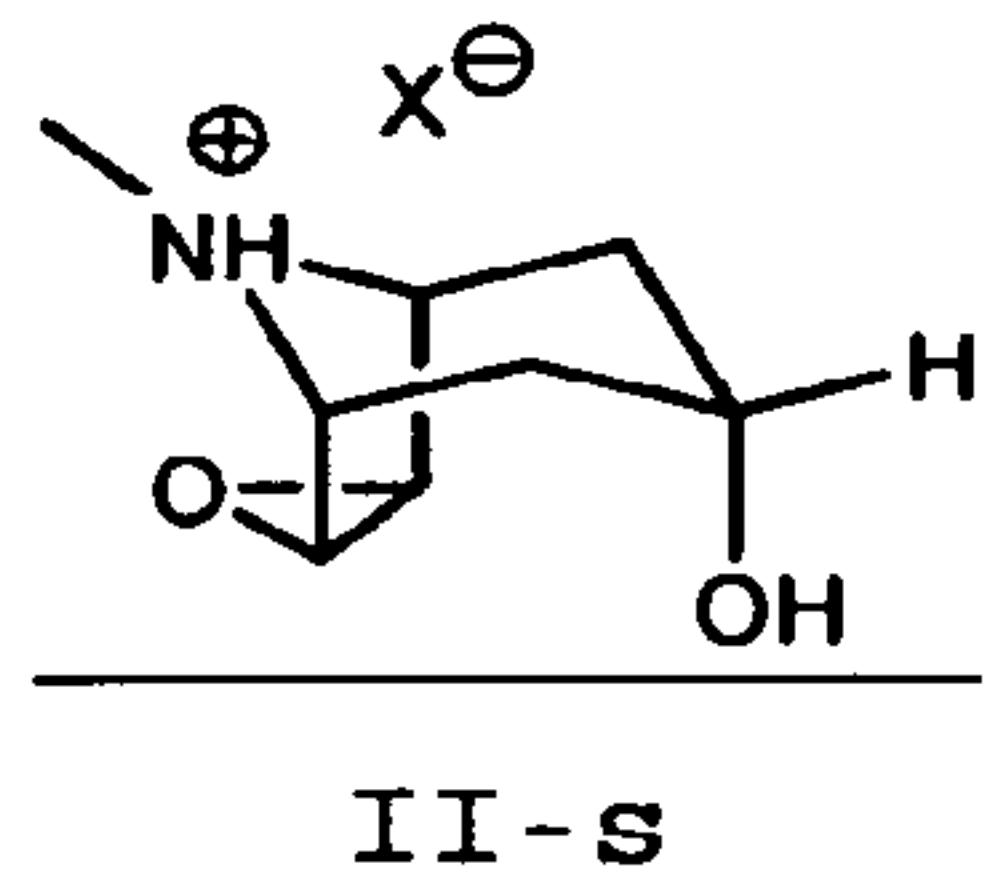
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comprising combining methyl-di-(2-thienyl)-glycolate of

formula I,

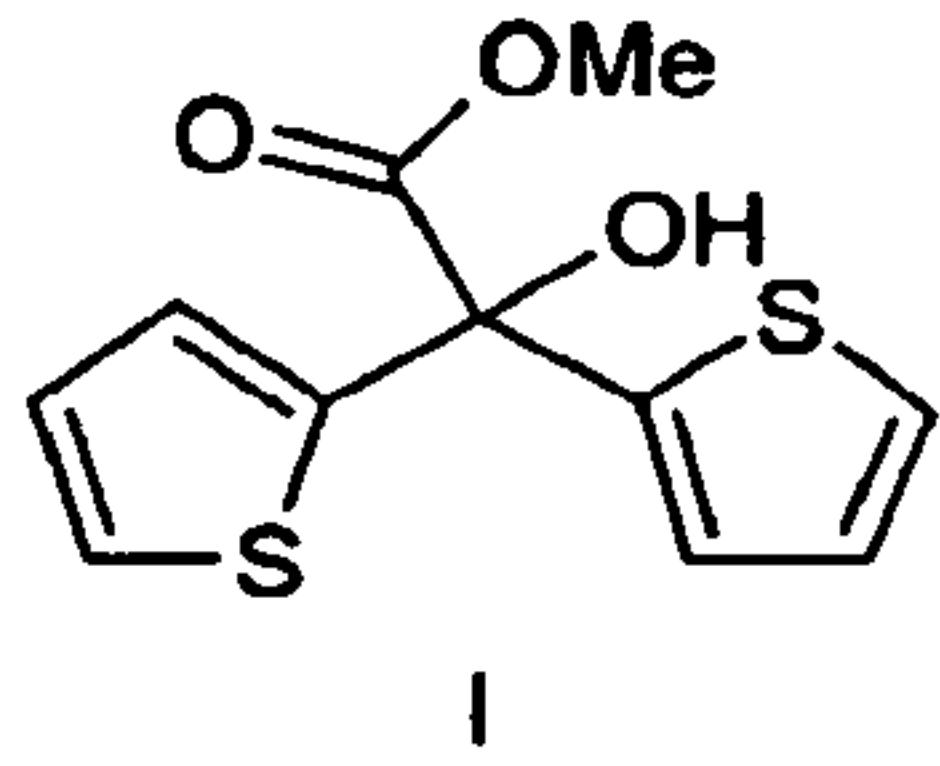


a weak inorganic base, a polar organic solvent and scopine salt of formula II-s



containing about 0.5% to about 40% by weight of salts to obtain a mixture, and heating the mixture.

[0015] In one embodiment, the present invention encompasses a process for the preparation of Tiotropium bromide comprising a) combining methyl-di-(2-thienyl)-glycolate of formula I,



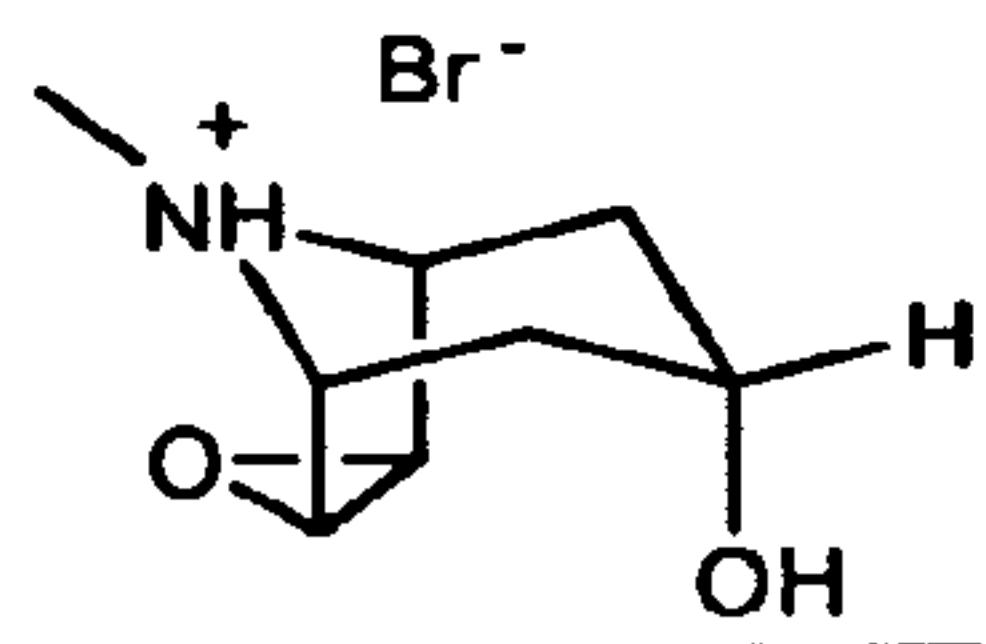
a weak base, a polar organic solvent, and scopine salt of formula II-s containing about 0.5% to about 40% by weight of salt to obtain a mixture; b) heating the mixture providing N-demethyl-tiotropium of formula III; c) recovering the N-demethyl-tiotropium of formula III; d) combining N-demethyl-tiotropium of formula III with methylbromide (MeBr) and an organic solvent to yield Tiotropium bromide.

[0016] In another embodiment, the present invention encompasses a process to prepare Tiotropium bromide, by preparing N-demethyl-tiotropium of formula III by the process of the present invention, and further converting it to Tiotropium bromide.

DETAILED DESCRIPTION OF THE INVENTION

[0017] As used herein, the term "room temperature" refers to a temperature ranging from about 20°C to about 26°C.

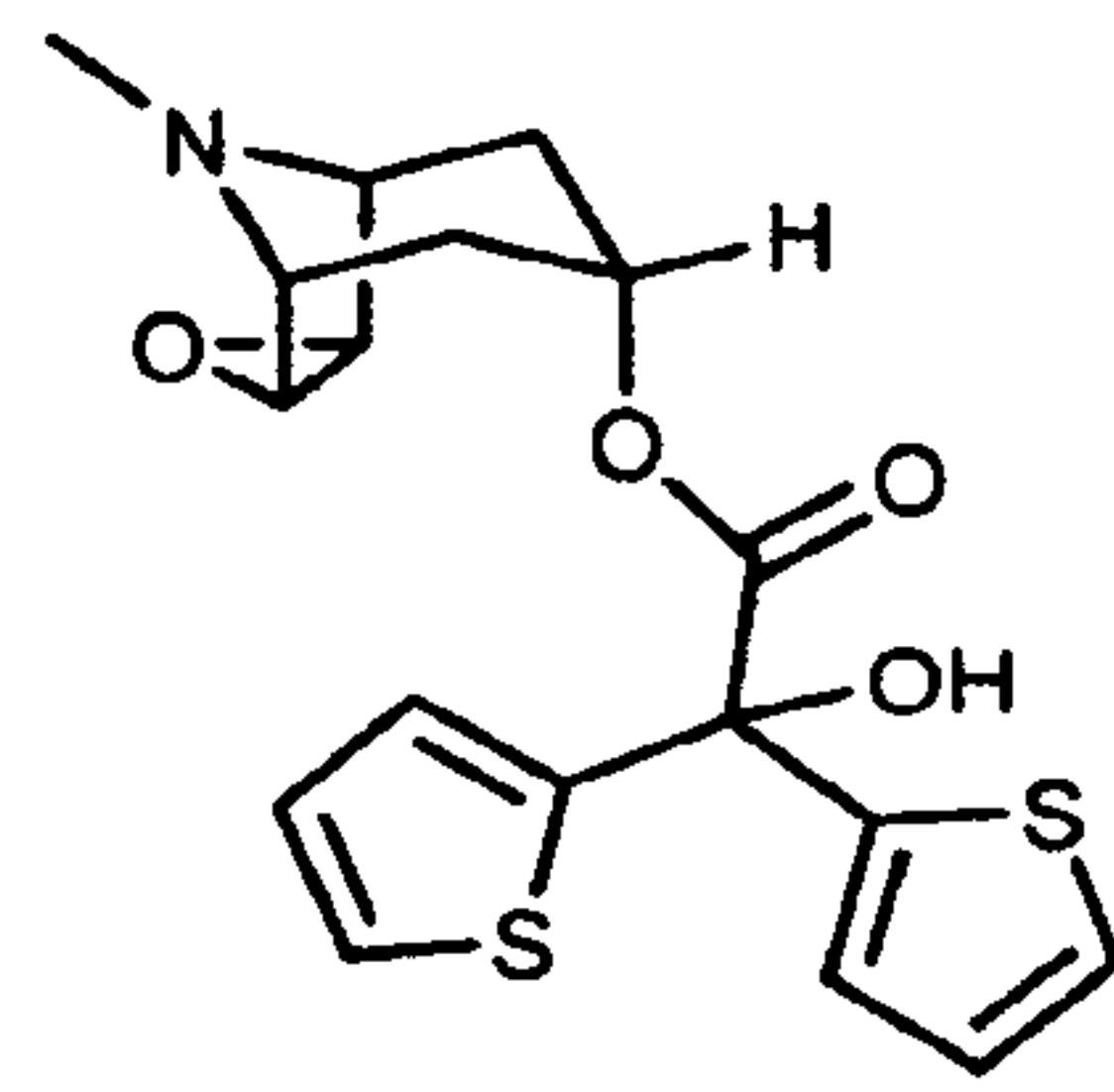
[0018] The present application discloses a novel approach to the synthesis of Tiotropium bromide, in particular from a Scopine hydrohalide of formula II-s, and even more specifically, from scopine HBr of the formula



that contains a low level of salts, as compared to the process disclosed in GB '781, where the obtained scopine HCl contained a high level of more than 60% by weight of inorganic salts. These salts are insoluble in water, and thus cannot be removed by a simple washing operation.

[0019] Typically, scopine hydrohalide containing low level of salts refers to scopine salt with less than 40% by weight of salts. Typically, the measurement of the salts content is done by sulfuric ashes, as exemplified in example 1.

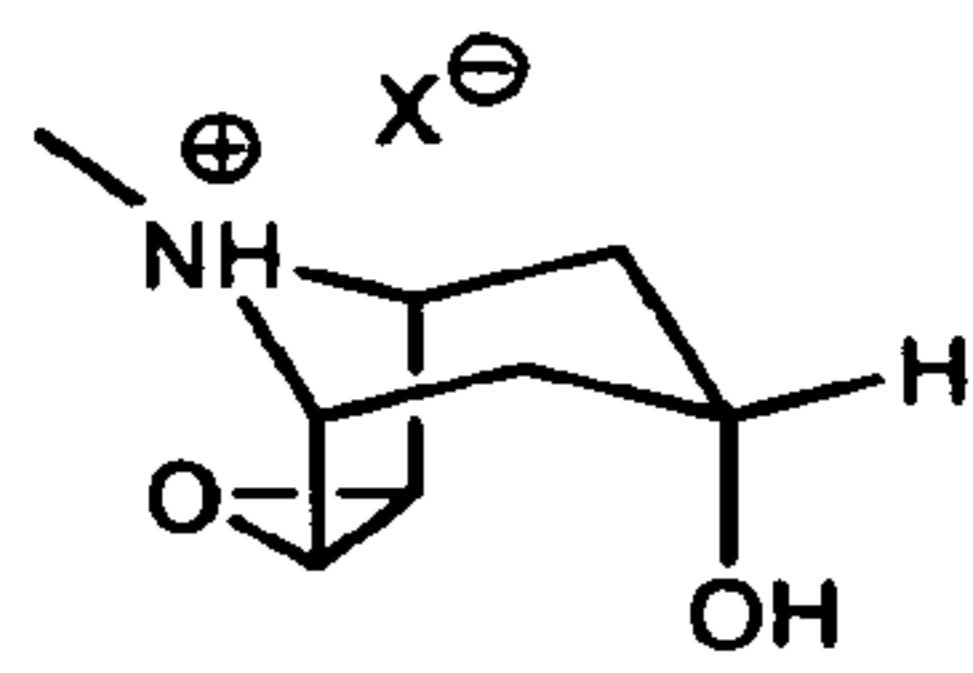
[0020] Moreover, the inventors of this application have found that the purity of the scopine salt affects the purity and the yield of N-demethyl-tiotropium of formula III, a key intermediate in the synthesis of Tiotropium bromide.



III

For example, without restricting the invention, using pure scopine salt leads to higher yield and purity, as will be further demonstrated.

[0021] The present invention encompasses scopine salt of formula II-s



II-s

containing about 0.5% to about 40% by weight of salts; wherein X is Br, Cl, SO₄, MeCOO, PO₄, MeSO₃, tartrate, fumarate, citrate, maleate, succinate, p-toluene sulphonate or amidosulphonate.

[0022] Preferably, scopine salt of formula II-s contains about 0.5% to about 20% by weight of salts, more preferably of about 0.5% to about 15% by weight, even more preferably of about 0.5% to about 5% by weight, and most preferably of about 0.5% to about 1% by weight. As mentioned before, the content of the salts can be determined by sulfuric ashes.

[0023] Preferably, the scopine salt is selected from a group consisting of HBr, HCl, H₂SO₄, CH₃COOH, H₃PO₄, MeSO₃H, tartrate, fumarate, citrate, maleate, succinate, malate, p-toluenesulphonate, and amidosulphonate. More, preferably, the salt is HBr.

[0024] The above scopine salt of formula II-s is obtained by filtration of the reaction mixture prior to the addition of acid. The process comprises: a) filtering solids from the reaction mixture containing scopine base and insoluble inorganic salts, such as borate salts, to produce a filtrate; b) washing the solids with a polar organic solvent; c) adding water to the filtrate to precipitate insoluble inorganic salts; d) filtering the filtrate to remove any precipitated salts; e) washing the inorganic salts with a polar organic solvent; and f) adding an acid and a polar organic solvent to the filtrate to obtain the scopine salt of formula II-s.

[0025] The addition of water in the above process causes the remaining insoluble salts to precipitate. Preferably, after the addition of water the filtrate is stirred for about 0.5 hour to about 2 hours, more preferably for about 1 to about 1.5 hours.

[0026] Typically, the addition of water results in a precipitate formation and thus a suspension. The suspension is concentrated under vacuum, to remove water, thus increasing the yield. Preferably, the suspension is concentrated at a temperature of no more than 55°C, more preferably of about 25°C to about 55°C, most preferably of about 30°C to about 35°C.

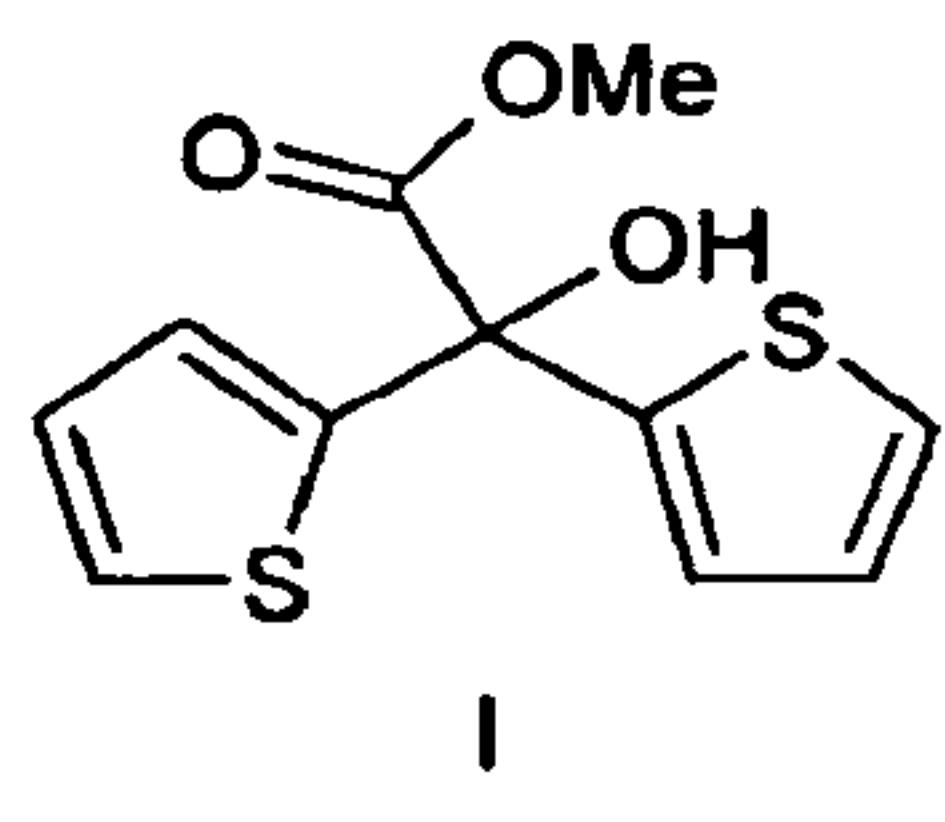
[0027] After the suspension is concentrated, it is filtered, and washed twice with a polar organic solvent. Preferably, the polar organic solvent is selected from a group consisting of C₁₋₆ alcohol, C₄₋₈ ether, C₃₋₁₀ ketone, C₂₋₄ nitrile, and mixtures thereof. Preferably, the C₁₋₆ alcohol is C₁₋₄ alcohol, more preferably C₁₋₃ alcohol. Preferably, the C₁₋₃ alcohol is methanol, ethanol or isopropanol. Moreover, absolute ethanol may be used. A preferred C₄₋₈ ether is C₄₋₆ ether, more preferably C₄₋₅ ether. A preferred C₄₋₅ ether is either tetrahydrofuran or

1,4-dioxane. Preferably, the C₃₋₁₀ ketone is C₂₋₃ ketone. Preferably, the C₂₋₃ ketones acetone. A preferred C₂₋₄ nitrile is C₁₋₂ nitrile. Preferably, the C₁₋₂ nitrile is acetonitrile. Most preferably, the solvent is ethanol.

[0028] Preferably, the acid is HBr.

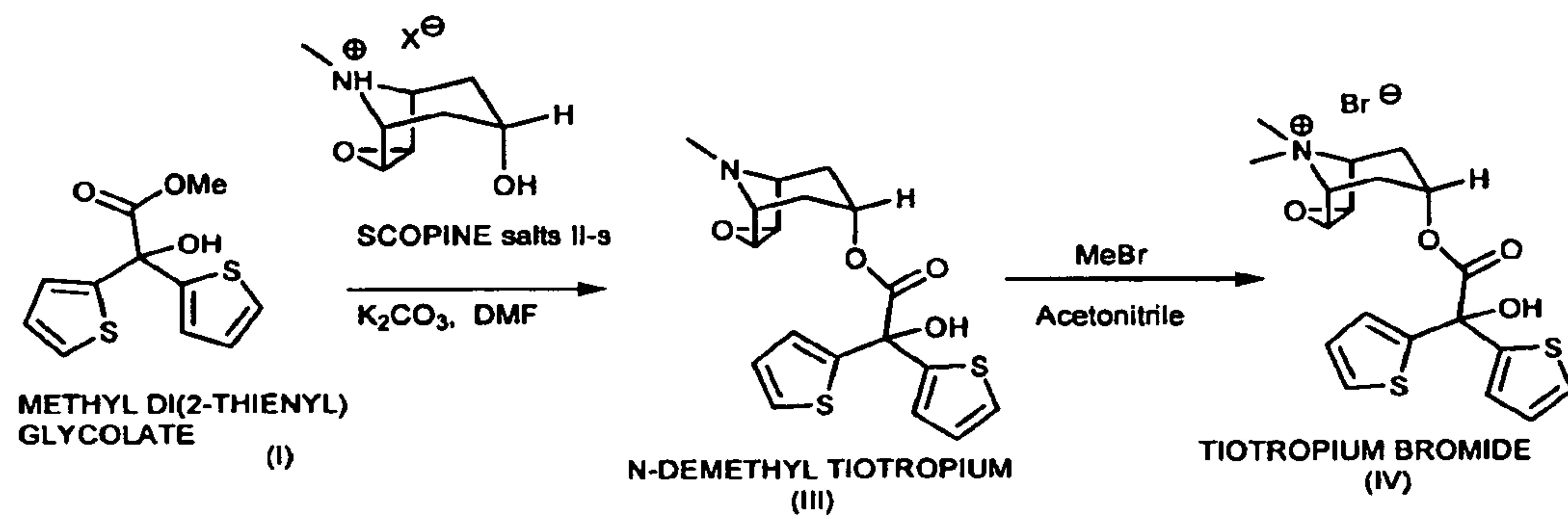
[0029] The obtained scopoline salt of formula II-s is then converted to Tiotropium bromide of formula IV. The conversion will be illustrated below.

[0030] The novel approach, also comprises a process for the preparation of Tiotropium bromide comprising: a) combining methyl-di-(2-thienyl)-glycolate of formula I,



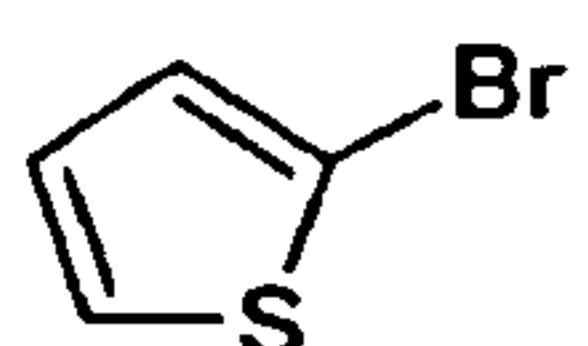
a weak inorganic base, a polar organic solvent, and scopoline salt of formula II-s containing about 0.5% to about 40% by weight of salt to obtain a mixture; b) heating the mixture providing N-demethyl-tiotropium of formula III; c) recovering the N-demethyl-tiotropium of formula III; d) combining N-demethyl-tiotropium of formula III with methylbromide (MeBr) and an organic solvent to yield Tiotropium bromide.

[0031] The process is illustrated by the following scheme:

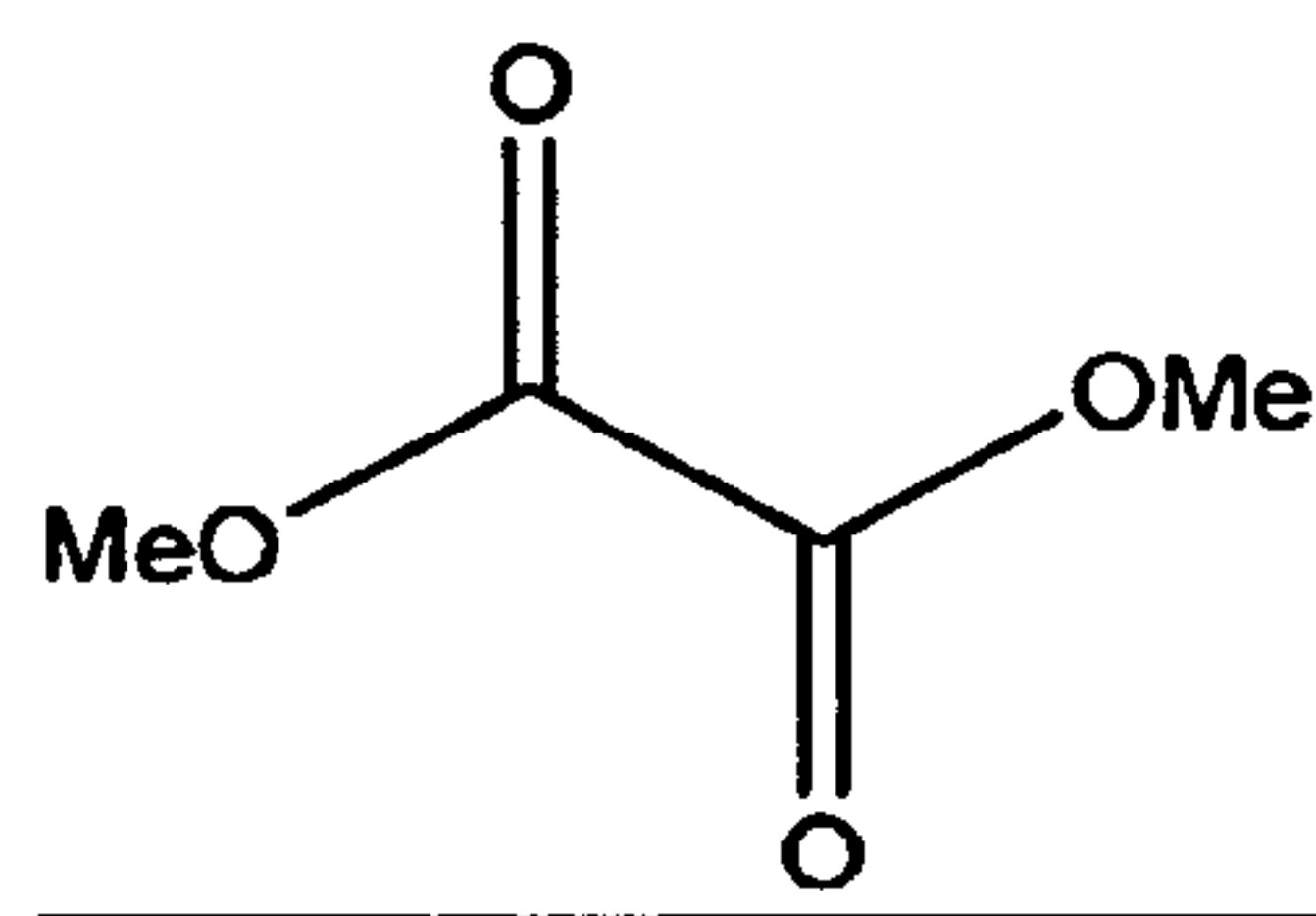


wherein X is Br, Cl, SO₄, MeCOO, PO₄, MeSO₃, tartrate, fumarate, citrate, maleate, succinate, p-toluene sulphonate or amidosulphonate.

[0032] The glycolate of formula I may be prepared by combining 2-bromo-thiophene of the following formula,



Mg, and an ethereal solvent; combining with dimethyloxalate of the following formula

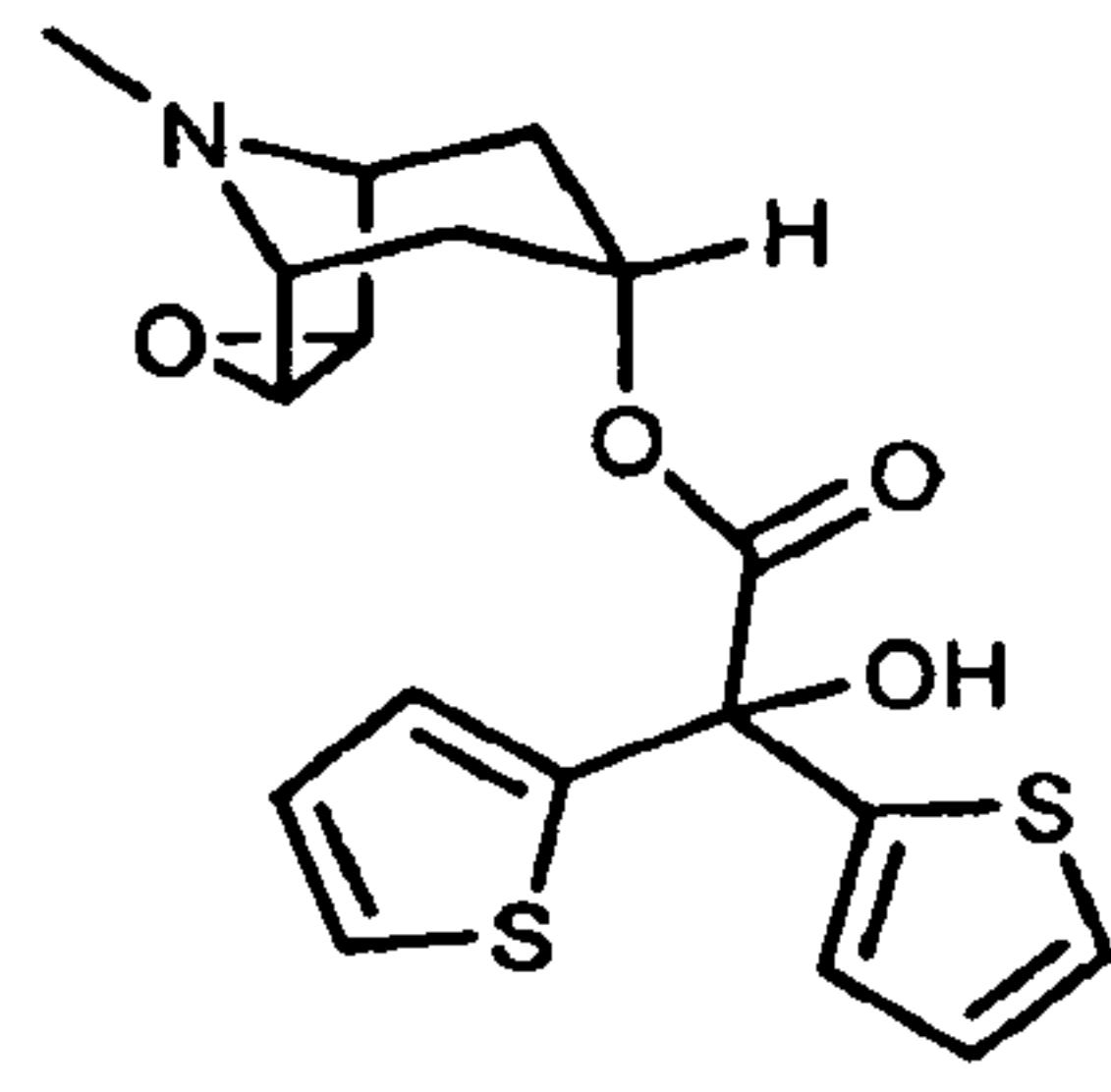


and quenching.

[0033] Combining 2-bromo-thiophene, Mg, and an ethereal solvent provides a Grignard reagent that can be prepared, for example, according to the process disclosed in Nyberg, K. *Acta Chemica Scandinavica*, 24, 1970, 1590-1596.

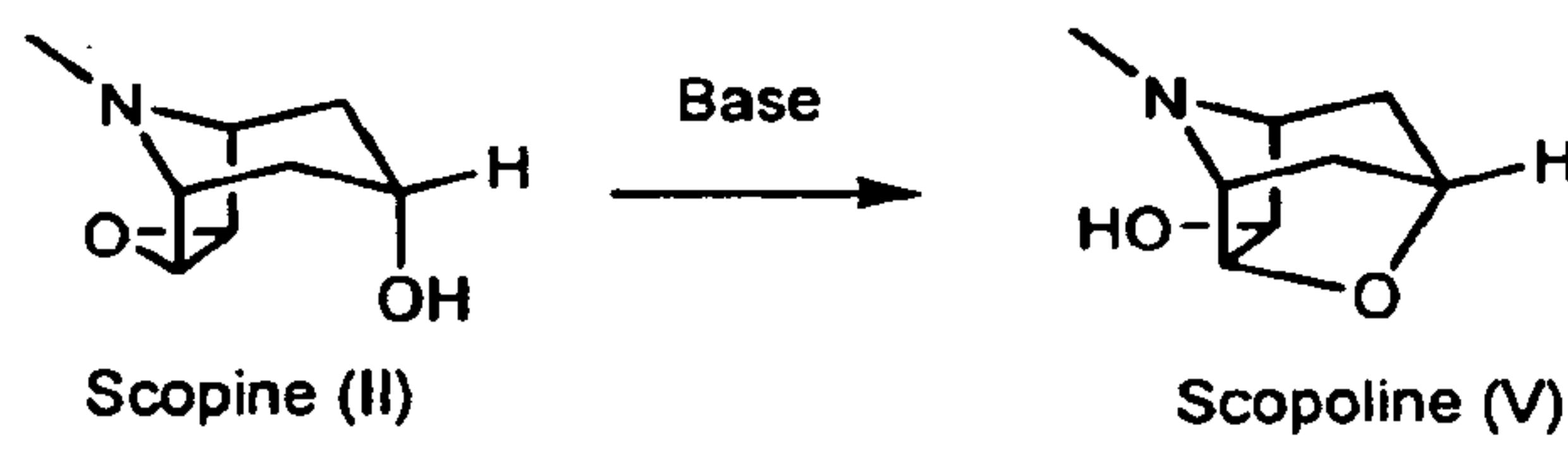
[0034] Methyl di-(2-thienyl)glycolate of formula I may be purified by crystallization from a mixture of ethanol and heptane, absolute ethanol and heptane, isopropanol and heptane, and from toluene and heptane.

[0035] N-demethyl-tiotropium of formula III

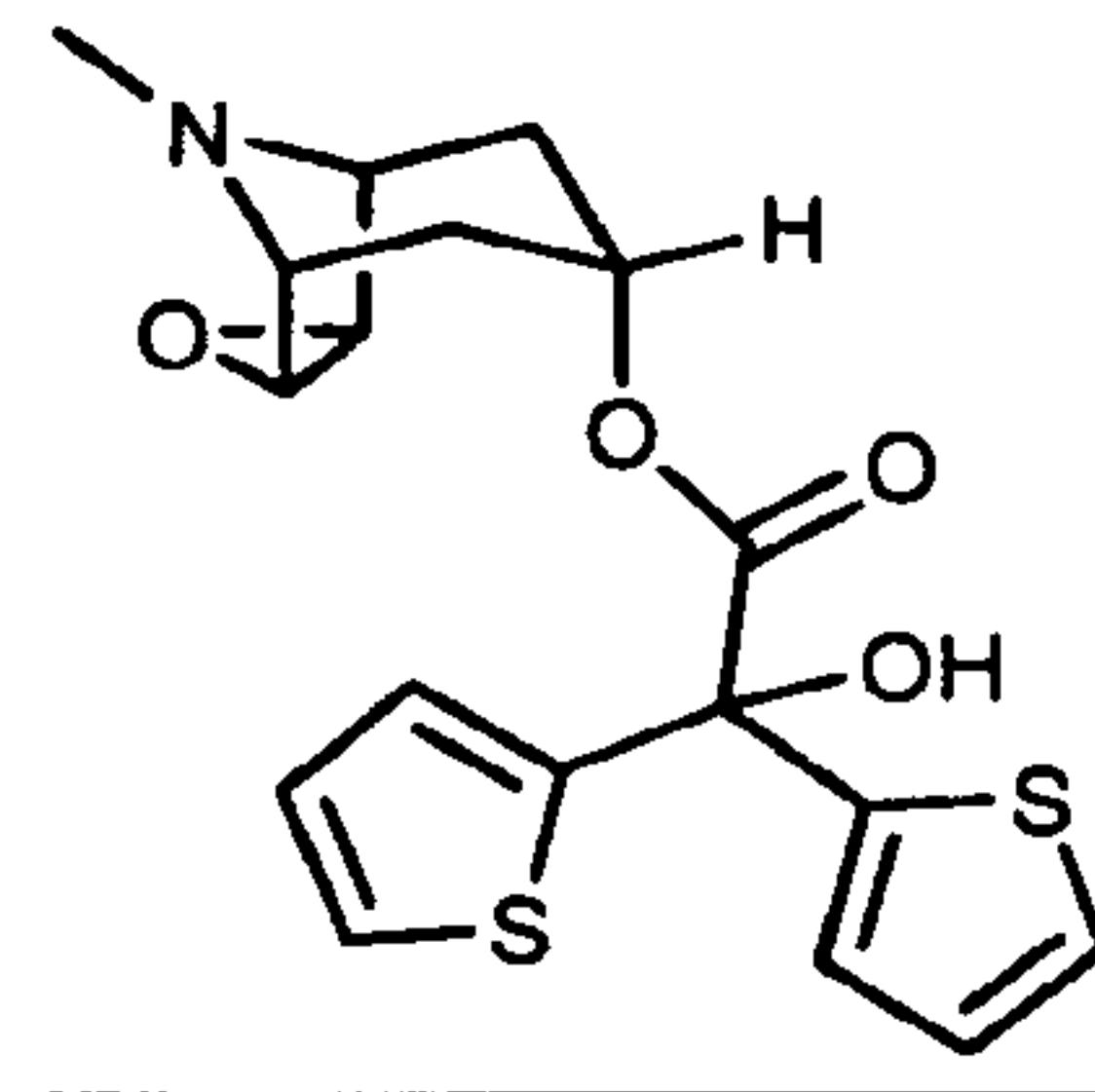


III

is prepared under conditions, which are not dangerous, are suitable for industrial scale preparation, and which provide a satisfactory yield. Moreover, the use of scopine salt instead of scopine base, and applying mild conditions in the preparation of N-demethyl-tiotropium of formula III, decrease significantly the conversion of scopine to scopoline, a side product that occurs in basic media, as illustrated in the following scheme:

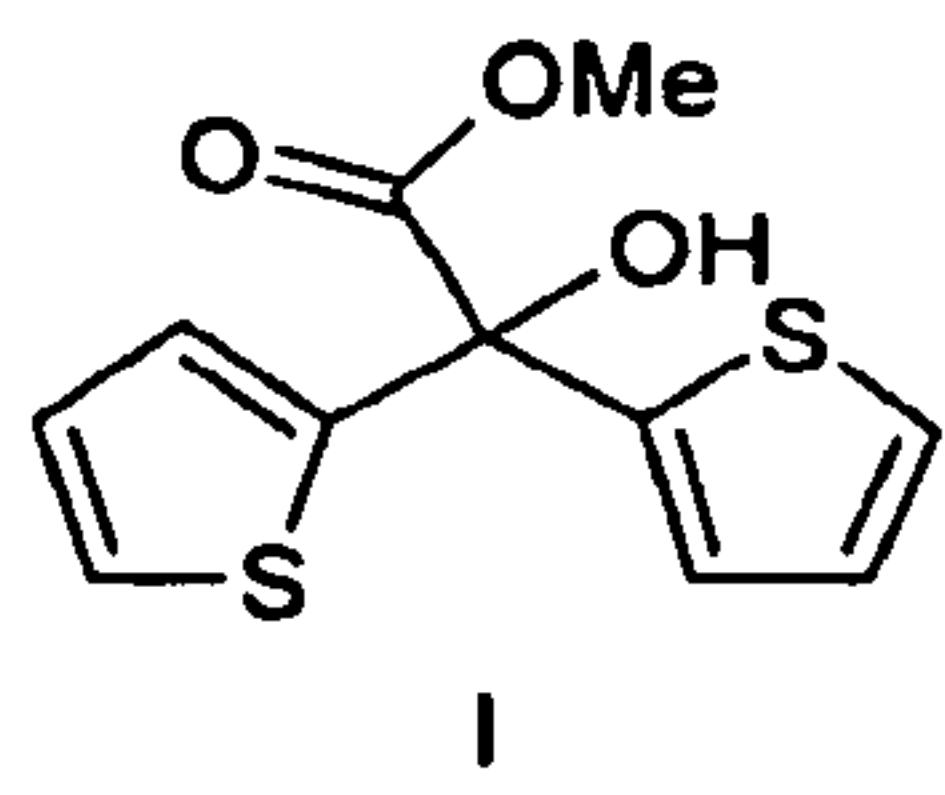


[0036] The preparation of N-demethyl-tiotropium of formula III,



III

comprises combining methyl-di-(2-thienyl)-glycolate of formula I,



a weak inorganic base, a polar organic solvent, and scopine salt of formula II-s, containing about 0.5% to about 40% by weight of salts to obtain a mixture, and heating the mixture.

[0037] Initially, scopine salt of formula II-s is suspended in a polar organic solvent. Preferably, the

salt is an HBr salt. The polar organic solvent is selected from a group consisting of amides, sulfoxides, sulfones, aromatic hydrocarbons, nitriles, and mixtures thereof. Preferably, the polar organic solvent is selected from the group consisting of C₁-C₄ amide, C₂-C₄ sulfoxide, C₂-C₄ sulfones, C₇-C₈ aromatic hydrocarbon, and C₂-C₄ nitrile. A preferred C₁-C₄ amide is dimethylformamide, N-methyl-2-pyrrolidone, or dimethylacetamide. A preferred C₂-C₄ sulfoxide is dimethylsulfoxide. Preferably, the C₂-C₄ sulfone is sulfolane. Preferably, the C₂-C₄ nitrile is acetonitrile. Preferably, the C₇-C₈ aromatic hydrocarbon is toluene. More preferably, the polar organic solvent is dimethylformamide.

[0038] Then, the weak inorganic base is added to the suspension providing a new suspension. Typically, an anhydrous weak inorganic base, i.e. having <0.5% of water by weight, is used in such reactions to obtain the free base form of scopine. Preferably, the weak inorganic base has a pKa of about 8 to about 12, even more preferably of about 9 to about 10. Preferably, the weak inorganic base is selected from a group consisting of: K₂CO₃, NaHCO₃, Na₂CO₃, Li₂CO₃, Cs₂CO₃, t-ButOK, and t-ButOLi. More preferably, the weak inorganic base is K₂CO₃.

[0039] After the addition of the base, a mixture of Methyl-di-(2-thienyl)-glycolate of formula I and another portion of the weak inorganic base are added to the new suspension, providing the mixture.

[0040] Methyl-di-(2-thienyl)-glycolate of formula I can be added as solid or in solution. Preferably, methyl-di-(2-thienyl)-glycolate of formula I is in solution in the polar organic solvent.

[0041] Preferably, the weak inorganic base is present in the mixture in an amount of about 0.45 to about 2.5,

more preferably, of about 2 to about 2.5 mole equivalent per mole equivalent of scopine salt of formula II-s. The weak inorganic base is added in two portions to prevent decomposition of scopine and to provide favorable reaction conditions.

[0042] Preferably, the mixture of the weak inorganic base and methyl-di-(2-thienyl)-glycolate of formula I are added at a temperature of about 25°C to about 65°C, more preferably at about 60°C to about 65°C.

[0043] Typically, the said mixture is heated in order to provide N-demethyl-tiotropium of formula III. Preferably, the said mixture is heated to a temperature of below 70°C, more preferably, of about 25°C to about 65°C, even more preferably of about 60°C to about 65°C, and most preferably of about 63°C to about 65°C. Preferably, heating is done for about 17 to about 24 hours, more preferably for about 18 to about 20 hours.

[0044] Preferably, heating is done under reduced pressure. Usually, the term "reduced pressure" refers to a pressure of about 70 to about 100 millibar.

[0045] Usually, such reactions are done under inert conditions, such as under an atmosphere of nitrogen. Inert conditions are provided by bubbling an inert gas, such as nitrogen and/or argon, during the reaction, through a second inlet. Preferably, nitrogen is bubbled in a rate of about 1.8 to about 2.6 L/min, more preferably, of about 2.0 to about 2.4 L/min, and even more preferably of about 2.2 to about 2.4 L/min.

[0046] The heating under reduced pressure, while bubbling nitrogen from a second inlet, assists in evaporating methanol, which is formed during the reaction, hence, shifting the reaction towards the formation of the product.

[0047] N-demethyl-tiotropium of formula III may be recovered by cooling the mixture; adding an acid to the cooled mixture providing a two phase system comprising of an organic and aqueous phases; extracting the aqueous phase with an organic solvent; adding a base to the aqueous phase to precipitate N-demethyl-tiotropium of formula III; filtering the precipitate N-demethyl-tiotropium of formula III; washing and drying the N-demethyl-tiotropium. Preferably, the acid is HBr.

[0048] Preferably, the heated mixture is cooled to a temperature of about 10°C to about -10°C, more preferably to about 5°C to about 0°C.

[0049] Preferably, the addition of the acid provides a pH of about 2 to about 3.5. Preferably, the organic solvent is toluene.

[0050] Preferably, the weak inorganic base is added at a temperature of about 0°C to about 5°C. Preferably, the weak inorganic base is selected from a group consisting of: K₂CO₃, NaHCO₃, Na₂CO₃, Li₂CO₃, Cs₂CO₃, t-ButOK, and t-ButOLi. Preferably, the weak inorganic base is K₂CO₃.

[0051] Preferably, the precipitate is washed with water to obtain pH of about 7.

[0052] Optionally, scopoline base may be used instead of scopoline salt. When scopoline base is used, typically, a smaller amount of the weak inorganic base is required. Preferably, about 1 to 1.5 mole equivalent of weak inorganic base per mole equivalent of scopoline base may be used.

[0053] The obtained N-demethyl-tiotropium of formula III may then be converted to Tiotropium bromide. The conversion can be done, for example, by the process disclosed in United States. Patent No. 5,610,163 or by the process of the present invention described below.

[0054] The conversion of N-demethyl-tiotropium of formula III to Tiotropium bromide can be done by reacting N-demethyl-tiotropium of formula III with methyl bromide in an organic solvent.

[0055] Initially, N-demethyl-tiotropium of formula III and the organic solvent are combined to obtain a suspension. Preferably, the organic solvent is selected from a group consisting of C₂₋₄ nitrile, C₄₋₈ linear or cyclic ether, mixtures of C₂₋₄ nitrile and C₄₋₈ linear or cyclic ether, mixtures of C₇₋₈ aromatic hydrocarbon and C₂₋₄ nitrile, and mixtures of C₂₋₄ nitrile and C₃₋₁₀ ketone. Preferably, the C₂₋₄ nitrile is acetonitrile. A preferred C₄₋₈ linear or cyclic ether is tetrahydrofuran. Preferably, a mixture of C₂₋₄ nitrile and C₄₋₈ linear or cyclic ether is that of acetonitrile and tetrahydrofuran. A preferred mixture of C₇₋₈ aromatic hydrocarbon and C₂₋₄ nitrile is that of toluene and acetonitrile. Preferably, a mixture of C₂₋₄ nitrile and C₃₋₁₀ ketone is that of acetone and acetonitrile. Most preferably, the solvent is acetonitrile.

[0056] Methyl bromide is then added to the suspension to provide a mixture. Methyl bromide can be used as a gas or in solution. Preferably, methyl bromide is used in solution, where the solvent is an organic solvent that is described above.

[0057] The mixture is maintained at a temperature of about 20°C to about 40°C. Preferably, the mixture is maintained at a temperature of about 20°C to about 25°C.

[0058] Typically, the mixture is maintained to allow the formation of Tiotropium bromide. Preferably, the mixture is maintained for about 12 to about 64 hours, more preferably for about 18 to about 22 hours.

[0059] Tiotropium bromide may then be recovered by any method known in the art, such as filtering and drying. Tiotropium bromide may then be purified by crystallization from ethanol. Preferably, crude Tiotropium bromide is dissolved in ethanol. More preferably, the ethanol is absolute ethanol. Preferably, the dissolution is done by heating to a temperature of about 75 to about 78°C. Typically, after dissolution, the solution is cooled to a temperature of about 22 to about 25°C to induce precipitation of Tiotropium bromide. Preferably, cooling is done over a period of about 6 to about 8 hours.

[0060] The precipitate is recovered from the suspension by filtration, washed with absolute ethanol, and dried.

[0061] Having described the invention with reference to certain preferred embodiments, other embodiments will become apparent to one skilled in the art from consideration of the specification. The invention is further defined by reference to the following examples describing in detail the preparation of the composition and methods of use of the invention. It will be apparent to those skilled in the art that many modifications, both to materials and methods, may be practiced without departing from the scope of the invention.

[0062] EXAMPLES

[0063] Sulfuric ashes methodology for determining the level of salts in scopine HBr:

[0064] 1g (exactly weighted) of scopine HBr was placed in a platinum melting pot. Sulphuric acid 96% was added, and then placed in an oven at 600°C until a constant weight was achieved. The melting pot, with the organic residue inside, was weighted to afford the

percent of inorganic salts present in the original material.

[0065] Example 1 : Comparative example: Preparation of scopine hydrochloride according to GB '781

[0066] 10.0g (22.84 mmol) of scopolamine hydrobromide trihydrate was suspended in 100 mL of absolute ethanol, and cooled to about 0°C. Sodium borohydride (4.0g, 105.7mmol) was then added portion-wise while maintaining the temperature at a maximum of 30°C. 4.8 mL of water was then added to the reaction mixture. After 3.5 hours, the reaction was completed and 50 mL of diethyl ether was then added. The reaction was then cooled to 0°C, and acidified with 2M hydrochloric acid in diethyl ether to a pH of about 2. The suspension was stirred at room temperature for 30 minutes and then filtered on GochP3. The white solid was dried at 45°C under vacuum for 4 hours, yielding 9g of product containing 79% of salts determined by sulphuric ashes.

[0067] Example 2: Preparation of Methyl di-(2-thienyl)glycolate

[0068] 1050 mL of tetrahydrofuran was loaded in a 2L round bottomed flask. 22.6g (0.93 mol) of magnesium turnings were then added, and the mixture was kept at 35°C, while catalytic bromoethane (200 mg, 1.84 mmol) was loaded. 150g (0.92 mol) of 2-bromothiophene was added dropwise, and after about 15% (13 ml) of reagent was exothermicity was observed. The temperature was maintained at a maximum of 50-55°C, and the remaining 2-bromothiophene was then added. At the end of the addition, the reaction mixture was heated to 65°C for 1.5 hours to 2 hours, and then cooled to 25°C.

[0069] The Grignard solution thus formed was added drop-wise, in about 2.5 hours to 3 hours, into a solution of dimethyl oxalate (54.3g, 0.46 mol) dissolved in 300 mL

of tetrahydrofuran, while maintaining the temperature at maximum 5-10°C via cooling bath.

[0070] The solution was kept under stirring for 0.5 hours to 1.0 hours at 5-10°C, and then saturated ammonium chloride (1400 mL of a mixture of 650g solid ammonium chloride and 2000 mL water) was added at 0°C, while monitoring the temperature (maximum 15-20°C). Then, 150 mL of water and 625 mL of toluene were added. The separated organic phase was washed with water (900 mL), and then with brine (900 mL).

[0071] To the organic phase 7.5g of charcoal was added, and the mixture was heated to 40°C and stirred at this temperature for one hour.

[0072] The mixture was then filtered on decalite pad, and washed three times with toluene (3 x 150 mL)

[0073] This solution was concentrated at 50-55°C under vacuum (about 30mmHg) to 270 mL, heated to 65°C, and then 720 mL n-heptane was added drop-wise over 2.5 hours.

[0074] The solution was stirred for one hour at 65°C, and then cooled in 3 hours to 25°C (about 1°C/5min). It was then left stirring at this temperature for at least 8 hours, filtered on GochP3, and washed once with 150mL n-heptane.

[0075] The creamy solid obtained was dried under vacuum at 45°C for 7 hours yielding 73.6g (63% overall yield, HPLC purity 99.8 area%).

[0076] Example 3: Crystallization of Methyl di-(2-thienyl)glycolate in Ethanol 96%/n-heptane

[0077] Crude methyl di-(2-thienyl)glycolate (2.0g) was dissolved in ethanol 96% (8.0 ml) at 45°C. 16.0 mL of n-heptane were then added drop-wise at 45°C in 20 minutes. The solution was maintained at 45°C for hour, and then it was cooled to 0°C in 1 hour, and left at this

temperature for another hour. The solid was filtered on a sintered glass funnel and it was washed once with n-heptane (2 mL). Drying for 6 hours at 50°C under vacuum yielded 1.4g of methyl di-(2-thienyl)glycolate (70%).

[0078] Example 4: Crystallization of Methyl di-(2-thienyl)glycolate in Absolute Ethanol /n-heptane

[0079] Crude methyl di-(2-thienyl)glycolate (10.0g) was dissolved in absolute ethanol (30.0 mL) at 55°C. 80.0 mL of n-heptane was then added drop-wise at 55°C in 30 minutes. The solution was maintained at 55°C for 1 hour, and then it was cooled to room temperature over 3 hours, and left at this temperature for 6 hours. The solid was filtered on a sintered glass funnel and it was washed once with n-heptane (10.0 mL). Drying for 18 hours at 50°C under vacuum yielded 8.0g of Methyl di-(2-thienyl)glycolate (80%).

[0080] Example 5: Crystallization of Methyl di-(2-thienyl)glycolate in Isopropanol /n-heptane

[0081] Crude methyl di-(2-thienyl)glycolate (5.0g) was dissolved in isopropanol (20.0 mL) at 60°C. 40.0 mL of n-heptane was then added drop-wise at 60°C in 30 minutes. The solution was maintained at 45°C for 1 hour, and then it was cooled to 0°C in 1 hour, and left at this temperature for another hour. The solid was filtered on a sintered glass funnel, and it was washed once with n-heptane (5.0 mL). Drying for 12 hours at 50°C under vacuum yielded 3.6g of methyl di-(2-thienyl)glycolate (72%).

[0082] Example 6: Crystallization of Methyl di-(2-thienyl)glycolate in Toluene /n-heptane

[0083] Crude methyl di-(2-thienyl)glycolate (250.0g) was dissolved in toluene (500 mL) at 55°C. 1750 mL of n-heptane was then added drop-wise at 55°C in 2.5 hours. The solution was maintained at 55°C for 1 hour, and then

it was left to cool to room temperature over 16 hours, and then cooled to 0°C, and left at this temperature for 2 hours. The solid was filtered on a sintered glass funnel and it was washed twice with n-heptane (2x150 mL). Drying for 16 hours at 50°C under vacuum yielded 184.1 of methyl di-(2-thienyl)glycolate (73.6%)

[0084] Example 7: Preparation of scopine hydrobromide in Ethanol 96%

[0085] 100g (0.228 mol) of scopolamine hydrobromide trihydrate was suspended in 1000 mL of ethanol 96%, and cooled to 0°C. Well ground sodium borohydride (30.23g, 0.80 mol) was then added portion-wise, while maintaining the temperature at maximum 25-30°C, gas evolution was noticed.

[0086] The reaction mixture was left stirring at room temperature for 16 hours, and then filtered on decalite pad, and washed twice with 100 mL of ethanol 96%. Water (19 mL) was added to the filtered solution, it was left stirring for 1.5 hours, and the obtained white suspension was concentrated under vacuum (about 30mmHg) at maximum 55°C to give a residue of 300 mL. After 30 minutes at room temperature, the residue was filtered on GochP3, and washed twice with 50 mL of ethanol 96%.

[0087] The solution was then cooled to 0°C, and acidified with 30 mL hydrobromic acid 48% to a pH of 1. 700 mL tetrahydrofuran was added drop-wise, in 3 hours, to the reaction mixture at 0°C, and it was maintained at 0°C for 5 hours. The white solid was then filtered on Goch P3, washed with 100 mL tetrahydrofuran, and dried at 50°C under vacuum for 16 hours, yielding 33.2g (64% yield, sulphuric ashes 0.7%).

[0088] Example 8: Preparation of scopine hydrobromide in abs.ethanol with precipitation in HBr 48%

[0089] 5g (11.4 mmol) of scopolamine hydrobromide trihydrate was suspended in 50mL of absolute ethanol, and cooled to 0°C. Well ground sodium borohydride (1.73g, 45.7mmol) was then added portion-wise while maintaining the temperature at a maximum of 30°C.

[0090] 4.0 ml water was then added to the reaction mixture, and then after 30 minutes, the suspension was filtered on decalite pad, and washed with 30mL of absolute ethanol.

[0091] The obtained solution was concentrated under vacuum at 40°C to residual 15mL and, after 2 hours at room temperature, filtered on GochP3, and washed with 4mL of absolute ethanol.

[0092] The solution was then cooled to 0°C, and acidified with hydrobromic acid 48% in water to obtain a pH of 1.

[0093] 15mL tetrahydrofuran was then added drop-wise to the reaction mixture at 0°C, and it was maintained at 0°C for 2 hours and then at room temperature for 2 days. The white solid was filtered on Goch P3, and dried at 50°C under vacuum for 16 hours, yielding 1.4g (52%). The other 40mL THF ("tetrahydrofuran") added to the mother liquor at 0°C led to obtain other 700mg product, containing 0.72% of salts.

[0094] Example 9: Preparation of scopine hydrobromide in abs.ethanol (5vol) with precipitation in HBr 48%

[0095] 10.0g (22.84 mmol) of scopolamine hydrobromide trihydrate was suspended in 50mL of absolute ethanol, and cooled to 0°C. Sodium borohydride (3.46g, 68.6mmol) was then added portion-wise while maintaining the temperature at a maximum of 30°C.

[0096] 4.8 ml of water was then added to the reaction mixture, and after 30 minutes the suspension was filtered

on decalite pad, and washed with 10mL of absolute ethanol.

[0097] The solution was then cooled to 0°C, and acidified with hydrobromic acid 48% in water to obtain a pH of 1.

[0098] 100mL tetrahydrofuran was added drop-wise to the reaction mixture at 0°C, and it was maintained at 0°C for 8 hours

[0099] The white solid was filtered on Goch P3, and dried at 50°C under vacuum for 16 hours, yielding 12.8g (52%), containing 0.65% of salts.

[0100] Example 10: Preparation of scopolamine hydrobromide in ethanol with precipitation in HBr 48%

[0101] 100g (0.22mol) of scopolamine hydrobromide trihydrate was suspended in 1000 mL of ethanol 1 96% and cooled to 0°C. Well ground sodium borohydride (30.23g, 0.80 mol) was then added portion-wise maintaining the temperature at maximum 25-30°C, gas evolution was noticed. The reaction mixture was left stirring at room temperature for 16 hours and then filtered on decalite pad, and washed twice with 100 mL of Ethanol 96%. 19 mL water was added to the filtered solution, and it was left stirring for 1.5 hours. The obtained white suspension was concentrated under vacuum (about 30mmHg) at maximum 55°C to residual 300 mL and, after 30 minutes at room temperature, filtered on GochGoch P3 and washed twice with 50 mL of Ethanol 96%. The solution was then cooled to 0°C and acidified with 30 mL hydrobromic acid 48% to a pH of 1. 1000 mL tetrahydrofuran was added drop-wise to the reaction mixture at 0°C in 3 hours and maintained at 0°C for 4-5 hours. The white solid was then filtered on Goch P3, washed with 100 mL tetrahydrofuran and dried at 50°C under vacuum for 16 hours, yielding 45.9g (85.6 % yield, sulphuric ashes 12.7%).

[0102] Example 11: Preparation of N-demethyl-tiotropium

[0103] 15g (0.064 mol) of scopine hydrobromide was suspended in 165 mL of dimethylformamide at 25°C, then 17.6g (0.127 mol) of anhydrous potassium carbonate were added, and the mixture was stirred at room temperature for about 60 minutes. 16.2g (0.064 mol) of methyl di-(2-thienyl)glycolate were dissolved in 30 mL of dimethylformamide, and, with 4.4 g (0.032 mol) of anhydrous potassium carbonate, they were added to the reaction mixture at about 60-65°C. The suspension was heated to 65°C, under vacuum (70-100 mbar), and under nitrogen stripping (2.2-2.4 L/min) for 18 hours. At the end of the reaction the distilled DMF ("dimethylformamide") was reintroduced to the reaction mixture and another 2 volumes of DMF, for a total of 15 volumes (225 mL), were added. The reaction mixture was cooled to 0°C, and acidified to pH 3 with about 168 mL of 2M HBr (the temperature during the addition below 20°C). The obtained solution was extracted twice with 85 mL toluene, and the combined aqueous phases were then cooled to 0-5°C, and basified with 8.5g of solid potassium carbonate to a pH of 9. After one hour at 0°C the solid was filtered on Goch P3, and washed five times with 60 mL of water to obtain a pH of 7. The solid was dried under vacuum at 45°C for 16 hours yielding 16.5g (69% yield, 98.3% HPLC purity)

[0104] Example 12: Preparation of N-demethyl-tiotropium with a scopine containing high percent of salts

[0105] 3.0g (0.012 mol) of scopine hydrobromide containing 69% inorganic salts was suspended at room temperature in 27 mL of dimethylformamide, then 3.4g (0.025 mol) of anhydrous potassium carbonate were added,

and the mixture was stirred at room temperature for about 60 minutes. 3.1g (0.012 mol) of methyl di-(2-thienyl)glycolate were dissolved in 9 mL of dimethylformamide, and, with 0.85 g (0.006 mol) of anhydrous potassium carbonate, they were added to the reaction mixture at about 60-65°C. The suspension was heated to 65°C, under vacuum (70-100 mbar), and under nitrogen stripping (2.2-2.4 L/min) for 18 hours. At the end of the reaction, the distilled DMF was re-added to the reaction mixture and other 2 volumes of DMF, for a total of 15 volumes (45 mL), were added. The reaction mixture was cooled to 0°C, and acidified to pH 3 with 33 mL of HBr 2M (temperature during addition below 20°C). The obtained solution was extracted twice with 15 mL toluene, and the combined aqueous phases were then cooled to 0-5°C, and basified with 7.1g of solid potassium carbonate to a pH of 9. After one hour at 0°C, the solid was filtered on Goch P3, and washed five times with 30 mL of water to obtain a pH of 7. The solid was dried under vacuum at 45°C for 16 hours yielding 1.8g (37.5% yield, 70% HPLC purity).

[0106] Example 13: Preparation of N-demethyl-tiotropium from scopoline base and a weak inorganic base

[0107] 5.45g (0.023mol) of scopoline hydrobromide was suspended in 30 mL of DCM, and then 5g (0.036 mol) of potassium carbonate was added. The reaction mixture was stirred at room temperature for one hour and then filtered on Goch P3 and washed with acetonitrile several times, (using about 10 ml of acetonitrile). After evaporation of the filtered solution, 2.78g of scopoline base was obtained (77.5 % yield). 3.0g (0.02 mol) of scopoline base was suspended at room temperature in 27 mL of dimethylformamide, then 2.7g (0.02 mol) of anhydrous potassium carbonate was added, and the mixture was

stirred at room temperature for about 60 minutes. 4.9g (0.02mol) of methyl di-(2-thienyl)glycolate were dissolved in 9 mL of dimethylformamide, and, with 1.33 g (0.0096 mol) of anhydrous potassium carbonate, they were added to the reaction mixture at about 60-65°C. The suspension was heated to 65°C, under vacuum (70-100 mbar), and under nitrogen stripping (2.2-2.4 L/min) for 18 hours. At the end of the reaction the distilled DMF was re-added to the reaction mixture and other 2 volumes of DMF, for a total of 15 volumes (45 mL), were added. The reaction mixture was cooled to 0°C, and acidified to a pH of 3 with 33 mL of 2M HBr (temperature during addition below 20°C). The obtained solution was extracted twice with 15 mL toluene, and the combined aqueous phases were then cooled to 0-5°C, and basified with 7.1g of solid potassium carbonate to a pH of 9. After one hour at 0°C, the solid was filtered on Goch P3, and washed five times with 30 mL of water to obtain a pH of 7. The solid was dried under vacuum at 45°C for 16 hours yielding 5.1g (69.5% yield, 98.5 % HPLC purity)

[0108] Example 14: Preparation of Tiotropium Bromide

[0109] 0.5g of N-demethyl tiotropium (1.33 mmol) was suspended in a flask under nitrogen with 5mL of CH₃CN. 0.525g of CH₃Br 48% w/w solution in CH₃CN (0.00266 mol) was loaded and the suspension was left under stirring at 22°C for 20 hours. The product was filtered and washed with 1mL of CH₃CN. 375 mg of Tiotropium was obtained (HPLC purity 92.21%, starting material 7.68%).

[0110] Example 15: Preparation of Tiotropium Bromide

[0111] 0.52g of N-demethyl tiotropium (1.39 mmol) was suspended in a flask under nitrogen with 5.23g of CH₃CN. 1.35g of CH₃Br 50% w/w solution in CH₃CN (0.0071 mol) was loaded, and the suspension was left under stirring at 22°C for 12 hours. The product was filtered and washed

with 1mL of CH₃CN. 572 mg of wet Tiotropium was obtained (HPLC purity 99.89%, starting material 0.07%).

[0112] Example 16: Preparation of Tiotropium Bromide

[0113] 4.96g of N-demethyl tiotropium (13.2 mmol) were loaded in a flask under nitrogen with 49.6mL of CH₃CN. A suspension was obtained. 12.61g of CH₃Br 50% w/w -CH₃CN solution- (0.066 mol) were loaded. The suspension was left under stirring at 22°C for 64 hours. The product was filtered and washed with 2mL of CH₃CN. 6.93g of wet Tiotropium was obtained, and dried under vacuum at 45°C for 22h (residual pressure 4 mbar). 5.9g of dry product (purity 99.8%, start 0.107%) was obtained.

[0114] Example 17: Crystallization of Tiotropium bromide

[0115] Crude Tiotropium bromide (1.00g) was dissolved in absolute ethanol (65 mL) at 78°C. The solution was heated to 78°C for about 30 min, and then was cooled to 22°C in at least 6 hours. The obtained suspension was maintained at 22°C for at least 3 hours, and then was filtered on a sintered glass funnel, and the solid was washed twice with absolute ethanol (2 x 1.0 mL). The solid was dried for 30 min. at 22°C under N₂ flow, and then for 9 hours at 60°C under reduced pressure (17 mbar). 0.66g of Tiotropium bromide was obtained.

[0116] Example 18: Preparation of Tiotropium bromide monohydrate from Tiotropium bromide ethanolate

[0117] 13.45g of dry Tiotropium bromide from example 16 was suspended in 80.7 mL of water and the suspension was stirred at room temperature for 4h. After it was filtered washing with 10 mL of water was conducted. The product was left on the filter under vacuum and under nitrogen at room temperature for 15 min. 11.66g of monohydrate was obtained. The content of water on the sample was 4.3% (TGA analysis).

[0118] Example 19: Comparative preparation of Tiotropium bromide from N-demethyltiotropium according to US 5,610,163

[0119] 10.0g (0.0265 mole) of scopine di-(2-thienyl)glycolate was dissolved in a mixture comprising 20 mL of anhydrous methylene chloride and 30 mL of anhydrous acetonitrile and treated with 12.8g (0.1325 mole) of methyl bromide (as 50% strength solution in anhydrous acetonitrile), and the reaction mixture was allowed to stand for 24 hours at room temperature in a tightly sealed reaction vessel. Crystals were precipitated during this time. They are filtered off under suction, washed using methylene chloride and dried at 35°C—under reduced pressure. White crystals were obtained (from methanol/acetone), m.p. 217.8 C. (decomposition) after drying at 111°C. under reduced pressure.

[0120] Example 20: Preparation of scopine hydrochloride in absolute ethanol (5vol) with precipitation in HCl 37%

[0121] 10.0g (22.84 mmol) of scopolamine hydrobromide trihydrate was suspended in 50mL of absolute ethanol, and cooled to 0°C. Sodium borohydride (3.46g, 68.6mmol) was then added portion-wise while maintaining the temperature at a maximum of 30°C.

[0122] 4.8 mL of water was then added to the reaction mixture, and after 30 minutes the suspension was filtered on decalite pad, and washed with 10mL of absolute ethanol.

[0123] The solution was then cooled to 0°C, and acidified with hydrochloric acid 37% in water to obtain a pH of 1.

[0124] 100mL of tetrahydrofuran was added drop-wise to the reaction mixture at 0°C, and it was maintained at 0°C for 8 hours

[0125] The white solid was filtered on Goch P3, and dried at 50°C under vacuum for 16 hours, yielding 12.8g (52%).

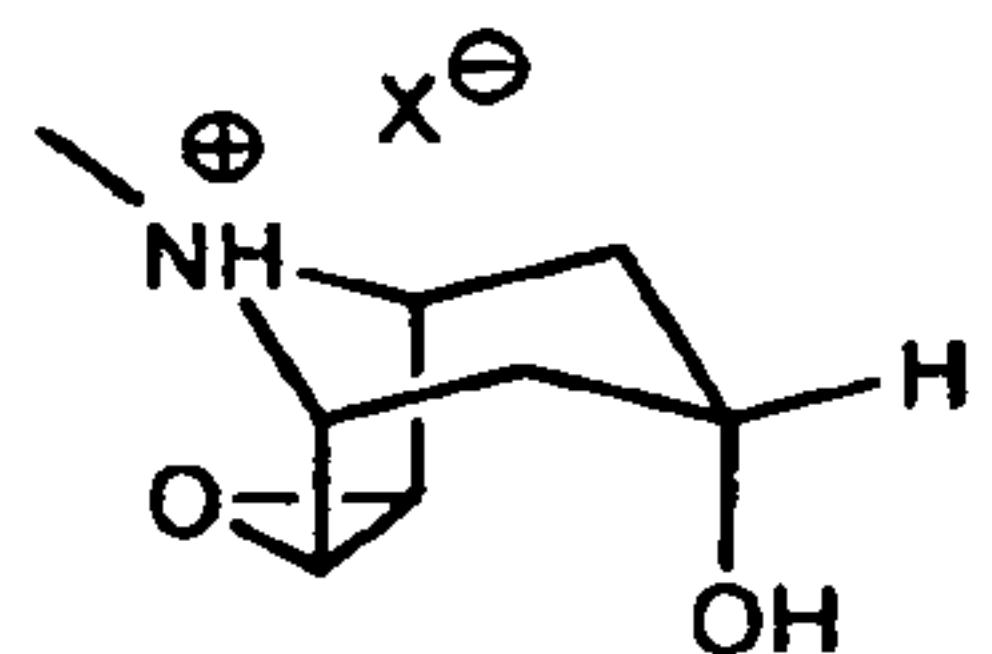
[0126] Example 21: Preparation of N-demethyl-tiotropium

[0127] 13.2g (0.064 mol) of scopoline hydrochloride was suspended in 165 mL of dimethylformamide at 25°C, then 17.6g (0.127 mol) of anhydrous potassium carbonate were added, and the mixture was stirred at room temperature for about 60 minutes. 16.2g (0.064 mol) of methyl di-(2-thienyl)glycolate was dissolved in 30 mL of dimethylformamide, and with 4.4 g (0.032 mol) of anhydrous potassium carbonate, they were added to the reaction mixture at about 60-65°C. The suspension was heated to 65°C, under vacuum (70-100 mbar), and under nitrogen stripped (2.2-2.4 L/min) for 18 hours.

[0128] At the end of the reaction, the distilled DMF was re-introduced to the reaction mixture and the other 2 volumes of DMF, for a total of 15 volumes (225 mL), were added. The reaction mixture was cooled to 0°C, and acidified to a pH of 3 with about 168 mL of 2M HBr (temperature during addition was below 20°C). The obtained solution was extracted twice with 85 mL of toluene, and the combined aqueous phases were then cooled to 0-5°C, and basified with 8.5g of solid potassium carbonate to a pH of 9. After one hour at 0°C, the solid was filtered on Goch P3, and washed five times with 60 mL of water to obtain a pH of 7. The solid was dried under vacuum at 45°C for 16 hours yielding 16.5g (69% yield, 98.3% HPLC purity).

What is claimed is:

1. A scopoline salt of formula II-s



Formula II-s

comprising about 0.5% to about 40% by weight of salts; wherein X is selected from the group consisting of Br, Cl, SO₄, MeCOO, PO₄, MeSO₃, tartrate, fumarate, citrate, maleate, succinate, p-toluene sulphonate and amidosulphonate.

2. The scopoline salt of claim 1, wherein X is Br.

3. The scopoline salt of claim 1 or claim 2, wherein said salt content is about 0.5% to about 20% by weight.

4. The scopoline salt of any of the preceding claims, wherein said scopoline salt is selected from the group consisting of an HBr, HC₁, H₂SO₄, CH₃COOH, H₃PO₄, MeSO₃H, tartrate, fumarate, citrate, maleate, succinate, malate, p-toluenesulphonate, borate and amidosulphonate scopoline salt.

5. A process of preparing said scopoline salt of any of the preceding claims, comprising:

- a. filtering a reaction mixture containing said scopoline salt to produce a filtrate;
- b. adding water to said filtrate to precipitate insoluble salts;
- c. filtering said insoluble salts from said filtrate;
- d. adding an acid to said filtrate to precipitate said scopoline salt; and
- e. collecting said precipitated scopoline salt.

6. The process of claim 5, wherein said acid is HBr.

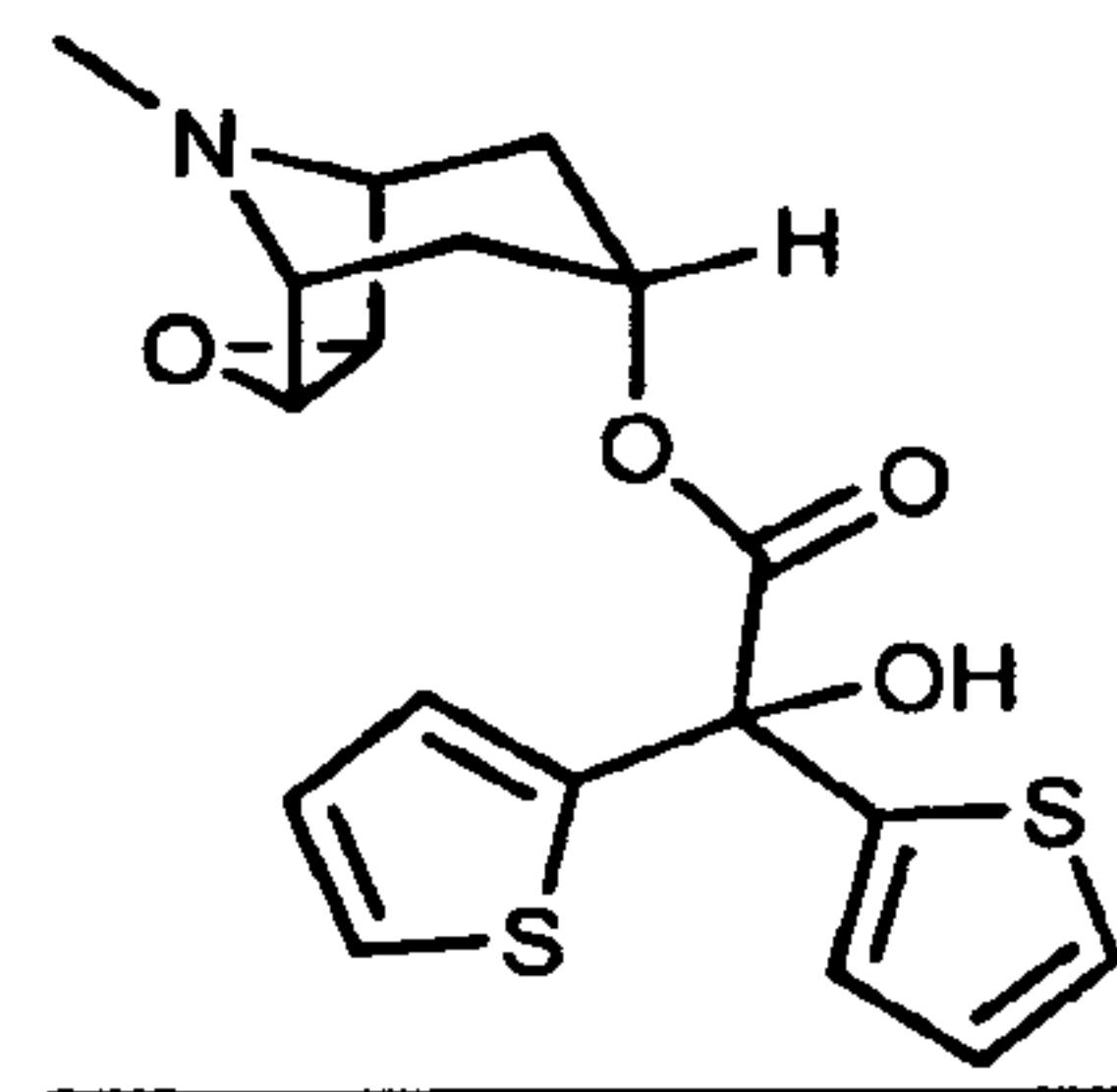
7. The process of claim 5 or claim 6, further comprising washing said precipitated scopine salts with a polar organic solvent.

8. The process of any of claims 5 to 7, wherein said polar organic solvent is selected from the group consisting of a C₁₋₆ alcohol, a C₄₋₈ ether, a C₃₋₁₀ ketone, a C₂₋₄ nitrile, and mixtures thereof.

9. The process of claim 8, wherein said polar organic solvent is selected from the group consisting of methanol, ethanol, isopropanol, 1,4-dioxane, acetone, acetonitrile, and mixtures thereof.

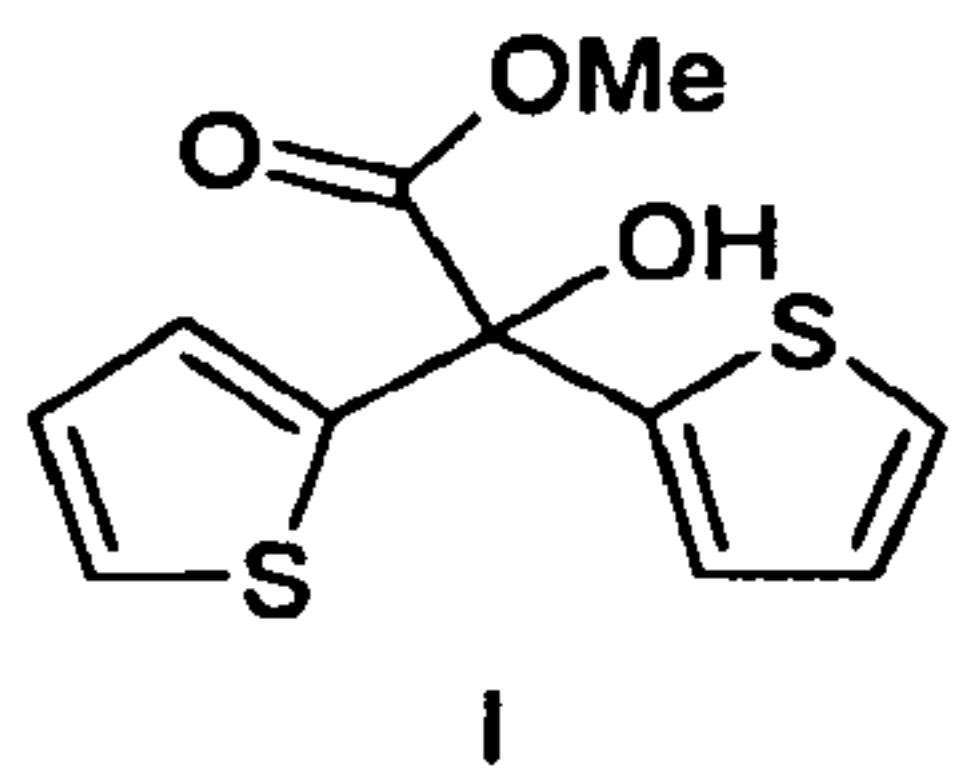
10. A process for preparing Tiotropium bromide comprising preparing scopine salt of formula II-s containing about 0.5% to about 40% by weight of salts by the process of claim 5, and converting it to Tiotropium bromide.

11. A process for the preparation of N-demethyl-tiotropium of formula III,



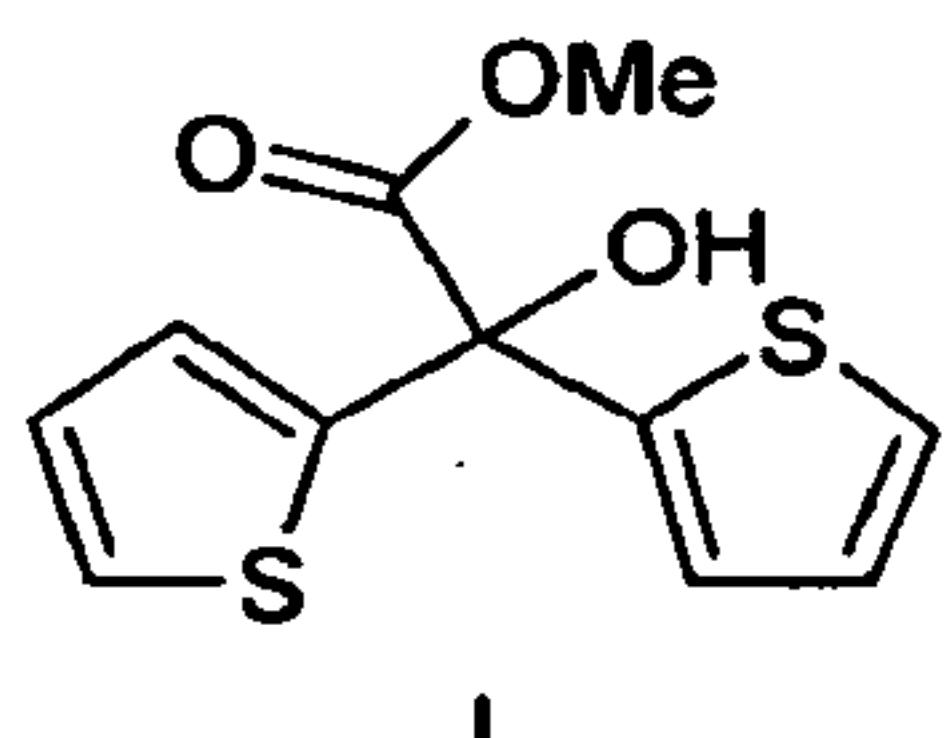
III

comprising reacting methyl-di-(2-thienyl)-glycolate of formula I,



with a scopine salt as defined in any of claims 1 to 4.

12. A process according to claim 11, wherein the process comprises combining methyl-di-(2-thienyl)-glycolate of formula I,



a weak inorganic base, a polar organic solvent, and a scropine salt of formula II-s containing about 0.5% to about 40% by weight of salts to obtain a mixture, and heating said mixture.

13. The process of any of claims 11 or 12, wherein said scropine salt is suspended in a polar organic solvent prior to said combination.

14. The process of claim 13, wherein said weak inorganic base is added to said suspension providing a new suspension.

15. The process of claim 14, wherein said weak inorganic base is anhydrous.

16. The process of any of claims 13 to 15, wherein said weak inorganic base has a pKa of about 8 to about 12.

17. The process of claim 14, wherein said weak inorganic base is selected from the group consisting of K_2CO_3 , $NaHCO_3$, Li_2CO_3 , Cs_2CO_3 , t-ButOK, and t-ButOLi.

18. The process of claim 17, wherein said weak inorganic base is K_2CO_3 .

19. The process of any of claims 12 to 18, wherein after said addition of said weak inorganic base, a mixture of said Methyl-di-(2-thienyl)-glycolate and

another portion of said weak inorganic base are added to said new suspension to provide a mixture.

20. The process of any of claims 12 to 19, wherein said Methyl-di-(2-thienyl)-glycolate is added in solution in a polar organic solvent.

21. The process of any of claims 12 to 20, wherein said polar organic solvent is selected from a group consisting of C₁-C₄ amides, C₂-C₄ sulfoxides, C₂-C₄ sulfones, C₇-C₈ aromatic hydrocarbons, and C₂-C₄ nitriles.

22. The process of claim 21, wherein said polar organic solvent is dimethylformamide.

23. The process of any of claims 12 to 22, wherein an amount of said weak inorganic base is about 0.45 to about 2.5 moles per mole equivalent of said scopoline salt.

24. The process of any of claims 14 to 23, wherein said addition of said weak inorganic base is performed at a temperature of about 25°C to about 65°C.

25. The process of claim 24, wherein said mixture is heated to a temperature of below about 70°C.

26. The process of any of claims 11 to 25, further comprising recovering said N-demethyl-tiotropium.

27. The process of any of the claims 11 or 12, wherein a scopoline base is used instead of said scopoline salt.

28. The process of claim 11, wherein a scopoline base is used and about 1 to 1.5 mole equivalents of said weak inorganic base is used per mole equivalent of said scopoline base.

29. A process for the preparation of Tiotropium bromide comprising the steps of:

a. preparing N-demethyl-tiotropium by a process as defined in any of claims 11 to 28; and

b. reacting said N-demethyl-tiotropium with methylbromide in an organic solvent to form Tiotropium bromide.

30. The process of claim 29, wherein said organic solvent is selected from the group consisting of a C₂₋₄ nitrile, a C₄₋₈ linear or cyclic ether, a mixture of a C₂₋₄ nitrile and a C₄₋₈ linear or cyclic ether, a mixture of a C₇₋₈ aromatic hydrocarbon and a C₂₋₄ nitrile, and a mixture of a C₂₋₄ nitrile and a C₃₋₁₀ ketone.

31. The process of claim 30, wherein said organic solvent is selected from the group consisting of acetonitrile, tetrahydrofuran, a mixture of acetonitrile and tetrahydrofuran, a mixture of toluene and acetonitrile, and a mixture of acetone and acetonitrile.

32. The process of claim 31, wherein said organic solvent is acetonitrile.

33. Use of a scopoline salt as defined in any of claims 1 to 5 in a process for the manufacture of N-demethyl tiotropium of Formula III or Tiotropium bromide.