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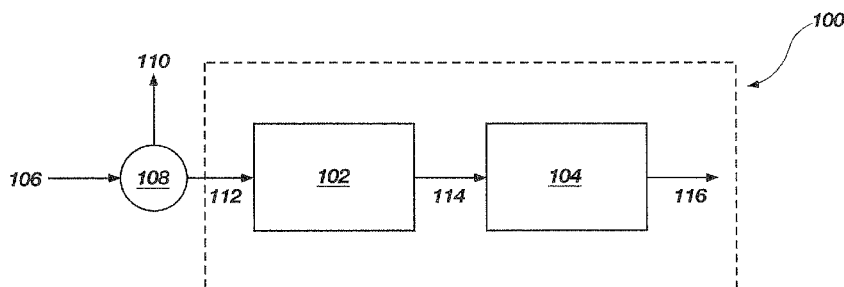


FIG. 1

(57) **Abstract:** A system for vaporizing and sublimating a slurry comprising a fluid including solid particles therein. The system includes a first heat exchanger configured to receive the fluid including solid particles and vaporize the fluid and a second heat exchanger configured to receive the vaporized fluid and solid particles and sublimate the solid particles. A method for vaporizing and sublimating a fluid including solid particles therein is also disclosed. The method includes feeding the fluid including solid particles to a first heat exchanger, vaporizing the fluid, feeding the vaporized fluid and solid particles to a second heat exchanger and sublimating the solid particles. In some embodiments the fluid including solid particles is liquid natural gas or methane including solid carbon dioxide particles.

TITLE OF THE INVENTION

SUBLIMATION SYSTEMS AND ASSOCIATED METHODS

5 RELATED APPLICATIONS

This application claims benefit of and priority to U.S. Non-provisional Patent Application Serial No. 12/938,967 filed November 3, 2010, SUBLIMATION SYSTEMS AND ASSOCIATED METHODS, which is incorporated herein by reference in its entirety.

10 The present application is related to co-pending U.S. Patent Application 11/855,071 filed on September 13, 2007, titled HEAT EXCHANGER AND ASSOCIATED METHODS, U.S. Patent Application 12/938,761 filed on November 3, 2010 and titled VAPORIZATION CHAMBERS AND ASSOCIATED METHODS, and copending U.S. Patent Application 12/938,826 filed on November 3, 2010 and titled HEAT EXCHANGER AND RELATED METHODS. The disclosure of each of the foregoing applications is incorporated herein by
15 reference in its entirety.

CONTRACTUAL ORIGIN OF THE INVENTION

This invention was made with government support under Contract Number DE-AC07-05ID14517 awarded by the United States Department of Energy. The government has
20 certain rights in the invention.

FIELD OF THE INVENTION

The invention relates generally to systems for vaporization and sublimation and methods associated with the use thereof. More specifically, embodiments of the invention
25 relate to a first heat exchanger configured to vaporize a fluid including solid particles therein and a second heat exchanger configured to sublimate the solid particles. Embodiments of the invention additionally relates to the methods of heat transfer between fluids, the sublimation of solid particles within a fluid, and the conveyance of fluids.

30 BACKGROUND

The production of liquefied natural gas is a refrigeration process that reduces the mostly methane (CH₄) gas to a liquid state. However, natural gas consists of a variety of gases in addition to methane. One of the gases contained in natural gas is carbon dioxide (CO₂). Carbon dioxide is found in quantities around 1% in most of the natural gas

infrastructure found in the United States, and in many places around the world the carbon content is much higher.

Carbon dioxide can cause problems in the process of natural gas liquefaction, as carbon dioxide has a freezing temperature that is higher than the liquefaction temperature of methane. The high freezing temperature of carbon dioxide relative to methane will result in solid carbon dioxide crystal formation as the natural gas cools. This problem makes it necessary to remove the carbon dioxide from the natural gas prior to the liquefaction process in traditional plants. The filtration equipment to separate the carbon dioxide from the natural gas prior to the liquefaction process may be large, may require significant amounts of energy to operate, and may be very expensive.

Small scale liquefaction systems have been developed and are becoming very popular. In most cases, these small plants are simply using a scaled down version of existing liquefaction and carbon dioxide separation processes. The Idaho National Laboratory has developed an innovative small scale liquefaction plant that eliminates the need for expensive, equipment intensive, pre-cleanup of the carbon dioxide. The carbon dioxide is processed with the natural gas stream, and during the liquefaction step the carbon dioxide is converted to a crystalline solid. The liquid/solid slurry is then transferred to a separation device which directs a clean liquid out of an overflow, and a carbon dioxide concentrated slurry out of an underflow.

The underflow slurry is then processed through a heat exchanger to sublime the carbon dioxide back into a gas. In theory this is a very simple step. However, the interaction between the solid carbon dioxide and liquid natural gas produces conditions that are very difficult to address with standard heat exchangers. In the liquid slurry, carbon dioxide is in a pure or almost pure sub-cooled state and is not soluble in the liquid. The carbon dioxide is heavy enough to quickly settle to the bottom of most flow regimes. As the settling occurs, piping and ports of the heat exchanger can become plugged as the quantity of carbon dioxide builds. In addition to collecting in undesirable locations, the carbon dioxide has a tendency to clump together making it even more difficult to flush through the system.

The ability to sublime the carbon dioxide back into a gas is contingent on getting the solids past the liquid phase of the gas and into a warmer section of a device without collecting and clumping into a plug. As the liquid natural gas is heated, it will remain at approximately a constant temperature of about -230° F (at 50 psig) until all the liquid has passed from a two-phase gas to a single-phase gas. The solid carbon dioxide will not begin to sublime back into a gas until the surrounding gas temperatures have reached approximately -

80° F. While the solid carbon dioxide is easily transported in the liquid methane, the ability to transport the solid carbon dioxide crystals to warmer parts of the heat exchanger is substantially diminished as liquid natural gas vaporizes. At a temperature when the moving, vaporized natural gas is the only way to transport the solid carbon dioxide crystals, the crystals may begin to clump together due to the tumbling interaction with each other, leading to the aforementioned plugging.

In addition to clumping, as the crystals reach warmer areas of the heat exchanger they begin to melt or sublime. If melting occurs, the surfaces of the crystals becomes sticky causing the crystals to have a tendency to stick to the walls of the heat exchanger, reducing the effectiveness of the heat exchanger and creating localized fouling. The localized fouling areas may cause the heat exchanger to become occluded and eventually plug if fluid velocities cannot dislodge the fouling.

In view of the shortcomings in the art, it would be advantageous to provide a system and associated methods that would enable the effective and efficient sublimation of solid particles found within a slurry. Additionally, it would be desirable for a system and associated methods to be able to effectively and efficiently warm and vaporize slurries of fluids containing solid particles.

BRIEF SUMMARY

In accordance with one embodiment of the invention, a method for vaporizing and sublimating a fluid including solid particles is provided. The method includes feeding a slurry comprising solid particles suspended in a first liquid to a first heat exchanger, vaporizing the first fluid in the first heat exchanger to form a first gas, feeding the first gas and the solid particles to a second heat exchanger, and sublimating the solid particles in the second heat exchanger to form a second gas.

In accordance with another embodiment of the invention, a method is provided for continuously vaporizing a slurry of liquid methane and solid carbon dioxide particles. The method includes feeding the slurry of liquid methane and solid carbon dioxide particles to a first heat exchanger, vaporizing the liquid methane in the first heat exchanger to form a mixture of solid carbon dioxide particles and gaseous methane, feeding the mixture of solid carbon dioxide particles and gaseous methane to a second heat exchanger, and sublimating the solid carbon dioxide particles in the second heat exchanger.

In accordance with a further embodiment of the invention, a system for vaporizing and sublimating a fluid including solid particles is provided. The system includes a first heat

exchanger configured to receive the fluid including solid particles and to vaporize the fluid and a second heat exchanger configured to receive the vaporized fluid and solid particles and to sublime the solid particles.

5 BRIEF DESCRIPTION OF THE DRAWINGS

While the specification concludes with claims particularly pointing out and distinctly claiming that which is regarded as the present invention, advantages of this invention may be more readily ascertained from the following detailed description when read in conjunction with the accompanying drawings in which:

10 FIGs. 1 and 2 are simplified schematics of a system for continuously vaporizing a fluid including solid particles suspended therein according to particular embodiments of the invention.

DETAILED DESCRIPTION

Some of the illustrations presented herein are not meant to be actual views of any
15 particular material, device, or system, but are merely idealized representations which are employed to describe the present invention. Additionally, elements common between figures may retain the same numerical designation.

FIG. 1 illustrates a system 100 according to an embodiment of the present invention. It is noted that, while operation of embodiments of the present invention is described in terms
20 of the sublimation of carbon dioxide in the processing of natural gas, the present invention may be utilized for the sublimation, heating, cooling, and mixing of other fluids and for other processes, as will be appreciated and understood by those of ordinary skill in the art.

The term “fluid” as used herein means any substance that may be caused to flow through a conduit and includes but is not limited to gases, two-phase gases, liquids, gels,
25 plasmas, slurries, solid particles, and any combination thereof.

As shown in FIG. 1, system 100 may comprise a first heat exchanger referred to herein as a vaporization chamber 102 and a second heat exchanger referred to herein as a sublimation chamber 104. In one embodiment, a product stream 106 including a plurality of solid particles suspended in a first fluid may be sent to a separator 108 to remove a portion of
30 the first fluid from the solid particles to form a fluid product stream 110 and a slurry 112 comprising the solid particles and a remaining portion of the first fluid. The slurry 112 may then be fed to the vaporization chamber 102. Within the vaporization chamber 102, the remaining first fluid in the slurry 112 may be vaporized, forming a first gas and the solid

particles 114. The first gas and the solid particles 114 may then be fed to the sublimation chamber 104. Within the sublimation chamber 104, the solid particles sublime, forming a second gas which is combined with the first gas and exits the sublimation chamber 104 as an exit gas 116. In one embodiment, the first fluid may comprise liquid natural gas and the solid particles may comprise solid carbon dioxide crystals.

FIG. 2 illustrates a more detailed schematic of one embodiment of the system 100 of FIG.1. As shown in FIG. 2, the slurry 112 of the solid particles and the first fluid are fed to the vaporization chamber 102. The slurry 112 may be at a pressure above the saturation pressure of the first fluid to prevent vaporization of the first fluid before entering the vaporization chamber 102. A second fluid 118 may also be fed to the vaporization chamber 102. The slurry 112 may be fed to the vaporization chamber 102 at a first temperature and the second fluid 118 may be fed to the vaporization chamber 102 at a second temperature, the second temperature being higher than the first temperature. The second fluid 118 mixes with the slurry 112 in a mixer 120 within the vaporization chamber 102. Within the mixer 120, heat may be transferred from the second fluid 118 to the slurry 112 causing the first fluid in the slurry 112 to vaporize forming the first gas and solid particles 114. At least about 95% of the first fluid in the slurry 112 may be vaporized within the vaporization chamber 102.

The vaporization chamber 102 may be configured to vaporize the first fluid in the slurry 112 without altering the physical state of the solid particles within the slurry 112. One embodiment of such a vaporization chamber is described in detail in previously referenced U.S. Patent Application 12/938,761 entitled "Vaporization Chamber and Associated Methods," and filed on November 3, 2010. Briefly, the vaporization chamber 102 may include a first chamber 140 surrounding a second chamber, which may also be characterized as a mixer 120. The second fluid 118 enters the first chamber 140 of the vaporization chamber 102 and envelops the mixer 120. Heat may be transferred from the second fluid 118 to the mixer 120 heating an outer surface of the mixer 120. The second fluid 118 also enters the mixer 120 and mixes with the slurry 112 as shown in broken lines within the vaporization chamber 102. In some embodiments, the mixer 120 may comprise a plurality of ports (not shown) that allow the second fluid 118 to enter the mixer 120 and promotes mixing of the second fluid 118 and the slurry 112. In additional embodiments, a wall of the mixer 120 may comprise a porous material which allows a portion of the second fluid 118 to enter the mixer 120 through the porous wall. In some embodiments, another portion of the second fluid 118' may exit the first chamber 140 of the vaporization chamber 102 and be directed to the

sublimation chamber 104. Alternatively, in some embodiments, the portion of the second fluid 118' may be directed to the sublimation chamber 104 before entering the vaporization chamber 102 as shown in broken lines.

As shown in FIG. 2, the first gas and the solid particles 114 formed in the vaporization chamber 102 may be fed to the sublimation chamber 104. A portion of the second fluid 118' is also fed to the sublimation chamber 104. A temperature of the portion of the second fluid 118' may be higher than a temperature of the solid particles from the first gas and the solid particles 114. Heat may be transferred from the portion of the second fluid 118' to the solid particles in the sublimation chamber 104, causing the solid particles to sublime and forming the second gas which mixes with the first gas and the portion of the second fluid 118' and forms the exit gas 116.

The sublimation chamber 104 may be configured to sublime the solid particles in the first gas and the solid particles 114 without allowing the particles to melt and stick together, fouling the system 100. One example of such a sublimation chamber 104 is described in detail in previously referenced U.S. Patent Application No. 12/938,826, entitled "Heat Exchanger and Related Methods," and filed November 3, 2010. Briefly, the sublimation chamber 104 may include a first portion 134 and a second portion 136. The first gas and the solid particles 114 may be fed into the first portion 134 of the sublimation chamber 104, and the portion of the second fluid 118' may be fed into the second portion 136 of the sublimation chamber 104. A cone-shaped member 138 may separate the second portion 136 from the first portion 134. At an apex of the cone-shaped member 138 is an opening or a nozzle 132 for directing the portion the second fluid 118' from the second portion 136 to the first portion 134 of the sublimation chamber 104. The nozzle 132 may comprise, for example, a changeable orifice or valve which may be sized to achieve a column of the second fluid 118" having a desired velocity extending through the first portion 134 of the sublimation chamber 104.

Particles from the first gas and the solid particles 114 may be entrained and suspended within the column of the second fluid 118". As the particles are suspended in the column of the second fluid 118", the column of the second fluid 118" heats the particles and causes the particles to sublime, forming the second gas. The cone-shaped member 138 helps direct the solid particles into the column of the second fluid 118".

The system 100 may be controlled using at least one valve and at least one temperature sensor. For example, as shown in FIG. 2, a first valve 122 may be used to control the flow of the second fluid 118 into the vaporization chamber 102 and a second valve 124 may be used to

control the flow of the portion of the second fluid 118' into the sublimation chamber 104. In some embodiments, the second valve 124 may be omitted and the flow of the second fluid 118, 118' into the vaporization chamber 102 and the sublimation chamber 104 may be controlled by the first valve 122. Temperature sensors may be placed throughout the system 100. For example, a first temperature sensor 126 may be located to determine the temperature of the second fluid 118 before the second fluid 118 enters the vaporization chamber 102. A second temperature sensor 128 may be located to determine the temperature of the first gas and the solid particles 114. A third temperature sensor 130 may be used determine the temperature of the exit gas 116. The temperatures at the second temperature sensor 128 and the third temperature sensor 130 may be controlled by varying the flow rate of the second fluid 118, 118' using the first valve 122 and the second valve 124. For example, if the temperature at the second temperature sensor 128 is too low, the flow through the first valve 122 (while the second valve 124 remains constant) may be increased to provide more of the second fluid 118 into the vaporization chamber 102. Alternatively, if the temperature at the second temperature sensor 128 is too low, the flow through the second valve 124 may be reduced thereby increasing of the pressure of the second fluid 118 in the vaporization chamber 102 and increasing the flow 118 flow into the 120 mixer. If the temperature at the third temperature sensor 130 is too low or if the flow of the portion of the second fluid 118' is too low through the nozzle 132, the flow of the portion of the second fluid 118' through the second valve 124 may be increased. The above operation controls are exemplary only and additional control mechanisms and designs may be utilized, as known in the art. In some embodiments, the first valve 122 and the second valve 124 may be controlled via a computer. Alternatively, in some embodiments, the first valve 122 and the second valve 124 may be controlled manually.

In one embodiment, the system 100 may be used as part of a liquefaction process for natural gas. For example, the present invention may be used in conjunction with an apparatus for the liquefaction of natural gas and methods relating to the same, such as is described in U.S. Patent No. 6,962,061 to Wilding et al., the disclosure of which is incorporated herein in its entirety by reference. The methods of liquefaction of natural gas disclosed in the Wilding patent include cooling at least a portion of a mass of natural gas to form a slurry which comprises at least liquid natural gas and solid carbon dioxide. The slurry is flowed into a hydrocyclone (i.e., the separator 108 as shown in FIG. 1) and forms a thickened slurry of solid carbon dioxide in liquid natural gas. The thickened slurry is discharged from the hydrocyclone through an underflow while the remaining portion of the liquid natural gas is flowed through an overflow of the hydrocyclone.

In this embodiment of the invention, the slurry 112 comprises a continuous flow of liquid natural gas and solid carbon dioxide particles as might be produced in a method according to the Wilding patent, as it is conveyed into the vaporization chamber 102. As the slurry 112 enters the mixer 120 within the vaporization chamber 102, the second fluid 118, which comprises a continuous flow of heated gas in this example (such as heated natural gas or heated methane), enters the vaporization chamber 102. The second fluid 118 heats the outside of mixer 120 and also enters the mixer 120, as desired. The heat from the second fluid 118 causes the liquid natural gas in the slurry 112 to vaporize. The temperature and pressure within the vaporization chamber 102 may be controlled such that the liquid natural gas in the slurry 112 vaporizes but that the solid carbon dioxide particles do not melt or sublimate. The second fluid 118 and the slurry 112 may be fed to the vaporization chamber 102 in about equal ratios. For example, in one embodiment, the mass flow rate of the second fluid 118 to the vaporization chamber 102 may be about one (1.0) to about one and a half (1.5) times greater than the mass flow rate of the slurry 112 to the vaporization chamber 102. In one embodiment, the mass flow rate of the second fluid 118 to the vaporization chamber 102 is about one and three tenths (1.3) times greater than the mass flow rate of the slurry 112 to the vaporization chamber.

As the slurry 112 is conveyed through the vaporization chamber 102, the initial heat energy provided by the second fluid 118 may be used to facilitate a phase change of the liquid methane of the slurry 112 to gaseous methane. As this transition occurs, the temperature of the slurry 112 may remain at about -230° F (this temperature may vary depending upon the pressure of the fluid) until all of the liquid methane of the slurry 112 is converted to gaseous methane. At this point, the solid carbon dioxide particles of the slurry 112 may now be suspended in the combined gaseous methane from the slurry 112 and second fluid 118 which exits the vaporization chamber 102 as the first gas and the solid particles 114. The temperature of the first gas and solid particles, determined by the second temperature sensor 128, may be controlled via the first valve 122 and the second valve 124 so that the temperature at the second temperature sensor 128 is higher than the vaporization temperature of the methane but colder than the sublimation temperature of the solid carbon dioxide particles. This ensures that the solid carbon dioxide particles do not begin to melt and become sticky within the vaporization chamber 102, preventing fouling of the vaporization chamber 102.

The first gas and the solid particles 114 comprising the vaporized methane and solid carbon dioxide particles are then continuously fed to the sublimation chamber 104. As the first gas and solid particles 114 enters the first portion 134 of the sublimation chamber 104, the portion of the second fluid 118', which again comprises a continuous flow of heated gas in this example (such as heated natural gas or heated methane), enters the second portion 136 of the sublimation chamber 104. The vaporized methane from the first gas and solid particles 114 exits the sublimation chamber 104 as part of the exit gas 116 while the solid carbon dioxide particles gather in the cone-shaped barrier 138. The portion of the second fluid 118' enters the first portion 134 of the sublimation chamber 104 through the nozzle 132 at about -80° F (this temperature may vary depending upon the pressure of the fluid environment) forming the column of the second fluid 118". The particles of carbon dioxide are funneled into the column of the second fluid 118" by the cone-shaped barrier 138 where the particles are suspended as they change phase from solid to vapor. All of the carbon dioxide particles may be converted to gaseous carbon dioxide. Once the gaseous carbon dioxide is formed, the gaseous carbon dioxide mixes with the gaseous methane from the first gas and solid particles 114 and the second fluid 118, 118' and exits the sublimation chamber as the exit gas 116.

The exit stream 116 may be monitored to maintain a temperature at the third temperature sensor 130 higher than the sublimation temperature of the solid carbon dioxide. However, it may be desirable to not overheat the exit stream 116 as the exit stream 116 may be reused as a refrigerant when cooling the natural gas to form the liquid natural gas according to the Wilding patent. In one embodiment, the temperature of the exit stream 116 may be maintained at about twenty degrees higher than the sublimation temperature of the solid carbon dioxide. For example, the exit stream 116 may be kept at about -40 °F and about 250 psia. By maintaining the exit stream 116 at about twenty degrees higher than the sublimation temperature of the solid carbon dioxide, all of the solid carbon dioxide in the exit stream 116 will be vaporized while still producing a cold stream for reuse in another heat exchanger.

In one example, the slurry 112 may enter the vaporization chamber 102 at about 245 psia and about -219° F at a mass flow rate of about 710 lbm/hr. The second fluid may enter the vaporization chamber 102 at about 250 psia and about 300° F at a mass flow rate of about 950 lbm/hr. The combined vaporized slurry, including the first fluid and the vaporized particles, and the second fluid may exit the system as the exit stream 116 at about -41°F and about 250 psia.

By using a separate vaporization chamber 102 and sublimation chamber 104 to form the exit gas 116, the process conditions (*i.e.*, pressure and temperature) for each of the vaporization chamber 102 and the sublimation chamber 104 may be optimized for gasifying the liquid and solid components of the slurry 112. By splitting the gasifying process of the slurry 112 into a vaporization chamber 102 and a sublimation chamber 104, the solid particles may be continuously sublimated without fouling the vaporization chamber 102. The system 100, therefore, provides a continuous method of transforming the slurry 112 into the exit gas 116, which may be easily disposed of.

In light of the above disclosure it will be appreciated that the apparatus and methods depicted and described herein enable the effective and efficient conveyance and sublimation of solid particles within a fluid. The invention may further be useful for a variety of applications other than the specific examples provided. For example, the described system and methods may be useful for the effective and efficient mixing, heating, cooling, and/or conveyance of fluids containing solids where there is a temperature difference between the vaporization temperature of the fluid and the sublimation temperature of the solid.

While the invention may be susceptible to various modifications and alternative forms, specific embodiments of which have been shown by way of example in the drawings and have been described in detail herein, it should be understood that the invention is not intended to be limited to the particular forms disclosed. Rather, the invention includes all modifications, equivalents, and alternatives falling within the scope of the invention as defined by the following appended claims and their legal equivalents.

CLAIMS

What is claimed is:

1. A method, comprising:
feeding a slurry comprising solid particles suspended in a first liquid to a first heat exchanger;
vaporizing the first fluid in the first heat exchanger to form a first gas;
feeding the first gas and the solid particles to a second heat exchanger; and
sublimating the solid particles in the second heat exchanger to form a second gas.
2. The method of claim 1, wherein feeding a slurry comprising solid particles suspended in a first liquid to a first heat exchanger comprises feeding a slurry comprising solid particles suspended in liquid natural gas to the first heat exchanger.
3. The method of claim 1, wherein feeding a slurry comprising solid particles suspended in a first liquid to a first heat exchanger comprises feeding a slurry comprising solid carbon dioxide particles suspended in a first fluid to the first heat exchanger.
4. The method of claim 1, further comprising feeding a second fluid to the first heat exchanger, the second fluid having a higher temperature than the slurry.
5. The method of claim 4, further comprising feeding a portion of the second fluid to the second heat exchanger.
6. The method of claim 1, wherein vaporizing the first fluid in the first heat exchanger to form a first gas comprises heating the slurry to a temperature higher than a vaporization temperature of the first liquid and lower than a sublimation temperature of the solid particles.
7. The method of claim 1, wherein vaporizing the first fluid in the first heat exchanger to form a first gas comprises:
feeding the slurry to a mixer;
filling a chamber around the mixer with a second, higher temperature, fluid to heat the mixer;
feeding a portion of the second, higher temperature fluid, into the mixer; and
mixing the slurry and the second, higher temperature, fluid to vaporize the first fluid.

8. The method of claim 1, wherein sublimating the solid particles in the second heat exchanger to form a second gas comprises:
feeding the first gas and solid particles to a first portion of the second heat exchanger;
feeding a second fluid to a second portion of the second heat exchanger;
supplying the second fluid from the second portion of the heat exchanger to the first portion of the heat exchanger and
sublimating the solid particles with heat from the second fluid.

9. The method of claim 8, wherein supplying the second fluid from the second portion of the heat exchanger to the first portion of the heat exchanger comprises supplying the second fluid from the second portion of the heat exchanger to the first portion of the heat exchanger through an opening formed in an apex of a cone shaped barrier member and into an interior portion of the cone shaped barrier member.

10. A method for continuously gasifying a slurry of liquid methane and solid carbon dioxide particles, comprising:
feeding a slurry of liquid methane and solid carbon dioxide particles to a first heat exchanger;
vaporizing the liquid methane in the first heat exchanger to form a mixture of solid carbon dioxide particles and gaseous methane;
feeding the mixture of solid carbon dioxide particles and gaseous methane to a second heat exchanger; and
sublimating the solid carbon dioxide particles in the second heat exchanger.

11. The method of claim 10, further comprising feeding additional gaseous methane to the first heat exchanger.

12. The method of claim 11, wherein vaporizing the liquid methane in the first heat exchanger to form a mixture of solid carbon dioxide particles and gaseous methane comprises transferring heat from the additional gaseous methane to the liquid methane to vaporize the liquid methane.

13. The method of claim 11, further comprising feeding a portion of the additional gaseous methane to the second heat exchanger.

14. The method of claim 13, wherein sublimating the solid carbon dioxide particles in the second heat exchanger comprises transferring heat from the portion of the additional gaseous methane to the solid carbon dioxide particles to sublime the solid carbon dioxide particles.

15. The method of claim 10, wherein vaporizing the liquid methane in the first heat exchanger to form a mixture of solid carbon dioxide particles and gaseous methane comprises vaporizing the liquid methane at a temperature lower than the sublimation temperature of the solid carbon dioxide particles.

16. A system for vaporizing and sublimating a slurry, comprising:
a first heat exchanger configured to receive the slurry comprising a fluid and solid particles and to vaporize the fluid; and
a second heat exchanger configured to receive the vaporized fluid and solid particles and to sublime the solid particles.

17. The system of claim 16, wherein at least one of the first heat exchanger and the second heat exchanger is configured to receive another fluid.

18. The system of claim 17, further comprising:
at least one temperature sensor configured to read a temperature of the vaporized fluid and solid particles; and
at least one valve configured to control the flow of the another fluid responsive to the temperature of the vaporized fluid and solid particles.

19. The system of claim 16, wherein the first heat exchanger comprises a chamber within a casing substantially surrounding a mixer.

20. The system of claim 19, wherein the mixer is configured to receive and mix the slurry and another fluid.

21. The system of claim 16, wherein the second heat exchanger comprises:
a first portion configured to receive the vaporized fluid and solid particles;
a second portion configured to receive another fluid; and

a cone-shaped barrier separating the first portion and the second portion, the cone-shaped barrier including an opening for transporting the another fluid into the first portion.

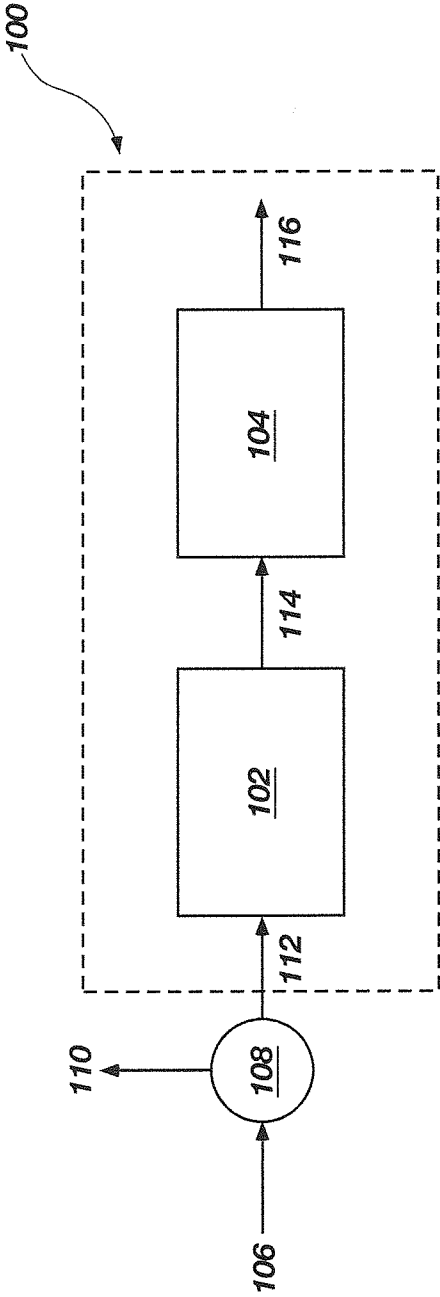


FIG. 1

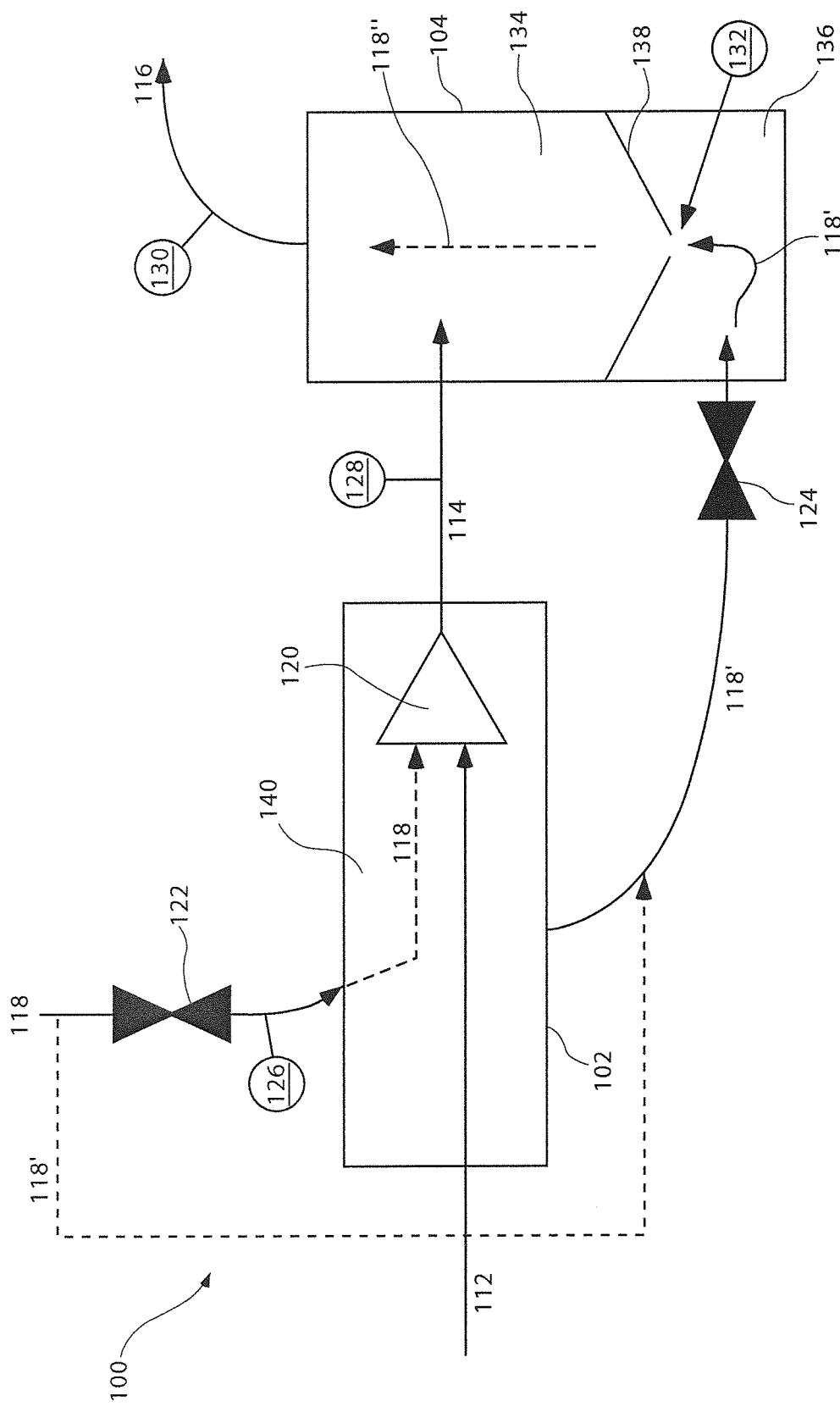


FIG. 2

INTERNATIONAL SEARCH REPORT

International application no.

PCT/US 11/59042

A. CLASSIFICATION OF SUBJECT MATTER

IPC(8) - F17C 11/00; C07C 9/00 (2012.01)

USPC - 62/46.1; 585/15; 62/46.2

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

USPC - 62/46.1; 585/15; 62/46.2

IPC(8) - F17C 11/00; C07C 9/00 (2012.01)

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

IPC(8): F17C 11/00; C07C 9/00

USPC: 62/46.1; 62/46.2; 585/15; search term delimited Patents, Non-Patent Literature.

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

Google Scholar; FPO; USPTO PubWEST.

Search terms used: carbon dioxide, evaporate, gas, gaseous, heat exchanger, methane, natural gas, particle, particulate, slurry, solid, sublimate, sublime, vapor s

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US 2009/0071634 A1 (Turner et al.) 19 March 2009 (19.03.2009), para [0035]-[0039], [0045]; Fig. 2-6	1-21
Y	US 2004/0177646 A1 (Wilkinson et al.) 16 September 2004 (16.09.2004), para [0001], [0073]-[0079]	1-21
A	US 3,724,225 A (Mancini et al.) 03 April 1973 (03.04.1973), col 2-4; Fig. 1, 2	1-21

☐ Further documents are listed in the continuation of Box C.


* Special categories of cited documents:

"A" document defining the general state of the art which is not considered to be of particular relevance

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"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

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Date of the actual completion of the international search

05 March 2012 (05.03.2012)

Date of mailing of the international search report

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