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(54) Title: METAL ION CATALYSIS OF METAL ION REDUCTION, METHODS, COMPOSITIONS, AND ARTICLES

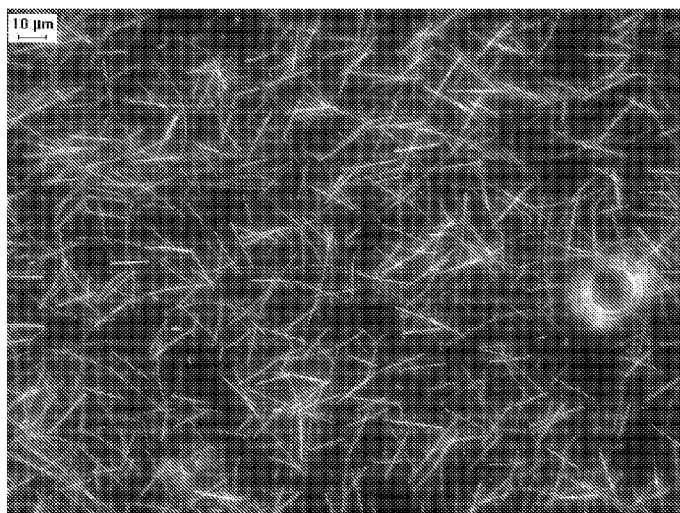


FIG. 1

(57) Abstract: Nanowire preparation methods, compositions, and articles are disclosed. Such methods which reduce metal ions to metal nanowires in the presence of complexes comprising metal-metal bonds, are capable of producing long, narrow, nanowires useful for electronics and optical applications.



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METAL ION CATALYSIS OF METAL ION REDUCTION,
METHODS, COMPOSITIONS, AND ARTICLES

BACKGROUND

5 The general preparation of silver nanowires (10-200 aspect ratio) is known. See, for example, *Angew. Chem. Int. Ed.* 2009, 48, 60, Y. Xia, Y. Xiong, B. Lim, S. E. Skrabalak, which is hereby incorporated by reference in its entirety. Such preparations typically employ Fe^{2+} or Cu^{2+} ions to “catalyze” the wire formation over other morphologies. The controlled preparation of silver
10 nanowires having the desired lengths and widths, however, is not known. For example, the Fe^{2+} produces a wide variety of lengths or thicknesses and the Cu^{2+} produces wires that are too thick for many applications.

 When iron or copper are used, they are typically provided as the metal halide salts FeCl_2 or CuCl_2 . See, for example, B. Wiley et al., *Nano Letters*,
15 **2004**, 4, 1733-1739 and K.E. Korte et al., *J. Mats. Chem.*, **2008**, 18, 437. Other metal halide salts have been used in nanowire synthesis. See, for example, J. Jiu, K. Murai, D. Kim, K. Kim, K. Suganuma, *Mat. Chem. & Phys.*, **2009**, 114, 333, which refers to NaCl , CoCl_2 , CuCl_2 , NiCl_2 and ZnCl_2 , and S. Nandikonda, “Microwave Assisted Synthesis of Silver Nanorods,” M.S. Thesis, Auburn
20 University, Auburn, Alabama, USA, August 9, 2010, which refers to NaCl , KCl , MgCl_2 , CaCl_2 , MnCl_2 , CuCl_2 , and FeCl_3 , and Japanese patent application publication 2009-155674, which discloses SnCl_4 . See also S. Murali et al., *Langmuir*, **2010**, 26(13), 11176-83; Z. C. Li et al., *Micro & Nano Letters*, **2011**, 6(2), 90-93; and B. J. Wiley et al., *Langmuir*, **2005**, 21, 8077. Japanese patent
25 application publication 2009-155674 discloses SnCl_4 .

SUMMARY

30 At least a first embodiment comprises methods comprising providing a composition comprising at least one first compound comprising at least one first reducible metal ion; and at least one second compound comprising at least a second metal atom and at least a third metal atom, said at least one second metal atom being bonded to the at least one third metal atom by at least

one metal-to-metal bond; and reducing the at least one first reducible metal ion to at least one first metal nanowire. In at least some cases, the composition may further comprise at least one solvent.

5 In at least some embodiments, the at least one first reducible metal ion comprises at least one coinage metal ion or at least one ion of an element from IUPAC Group 11 or at least one silver ion.

In at least some cases, the at least one second metal atom comprises molybdenum, or the at least one third metal atom comprises molybdenum, or both.

10 In at least some embodiments, the at least one metal-to-metal bond comprises one or more of at least one triple bond or at least one quadruple bond, or the at least one metal-to-metal bond comprises at least one quadruple bond.

In at least some cases, the at least one second compound comprises at least one complex comprising the at least one second metal atom and the at least one third metal atom, the at least one complex further comprising at least one carbon atom bonded to the at least one second metal atom or the at least one third metal atom. An exemplary second compound is $(\text{Mo}(\text{OCOCH}_3)_2)_2$.

Other embodiments provide the at least one first metal nanowire produced according to such methods.

20 At least a second embodiment provides methods comprising providing a composition comprising at least one compound comprising at least one first reducible metal ion; at least one second compound comprising at least a second metal atom and at least a third metal atom, where the at least one second metal atom is bonded to the at least one third metal atom by at least one metal-to-metal bond, and at least one solvent; and reducing the at least one first reducible metal ion to at least one first metal. The at least one first metal ion may, for example, comprise at least one element from IUPAC Group 11, or at least one coinage metal ion, or at least one silver ion. The at least one second metal atom, or the at least one third metal atom, or both, may, for example, comprise molybdenum. The at least one metal-to-metal bond may, in some embodiments, comprise one or more of at least one double bond, or at least one triple bond, or at least one quadruple bond. Or the at least one metal-to-metal bond may, in other embodiments, comprise one or more of at least one triple bond or at least one quadruple bond. Or the at least one metal-to-metal bond may, in still other

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embodiments, comprise at least one quadruple bond. The at least one second compound may, in some cases, comprise at least one complex comprising the at least one second metal atom and the at least one third metal atom, where the at least one complex further comprises at least one carbon atom bonded to the at least one second metal atom or the at least one third metal atom, or, in other cases, the at least one complex further comprises at least one first carbon atom bonded to the at least one second metal atom and at least one second carbon atom bonded to the at least one third metal atom. The at least one second compound may, for example, comprise $(\text{Mo}(\text{OCOCH}_3)_2)_2$. In at least some embodiments, the at least one solvent may comprise at least one polyol. In other embodiments, the at least one solvent may comprise at least one of ethylene glycol, propylene glycol, glycerol, one or more sugars, or one or more carbohydrates. In some cases, the composition has a ratio of the total moles of the at least one second metal atom to the total moles of the at least one first reducible metal ion from about 0.0001 to about 0.1, and, optionally, the composition may further have a ratio of the total moles of the at least one third metal atom to the total moles of the at least one first reducible metal ion from about 0.0001 to about 0.1. The reduction may, for example, be carried out at one or more temperatures from about 120 °C to about 190 °C. In at least some such embodiments, the composition may further comprise at least one protecting agent, such as, for example, polyvinylpyrrolidinone.

Other embodiments provide the at least one first metal produced according to such methods.

Still other embodiments provide at least one article comprising the at least one first metal produced according to such methods. Such articles may, for example, comprise at least one first metal comprising one or more nanowires, nanocubes, nanorods, nanopyramids, nanotubes, or the like. Or such articles may, for example, comprise at least one first metal comprising objects having an average diameter of between about 10 nm and about 500 nm, or an aspect ratio from about 50 to about 10,000.

These embodiments and other variations and modifications may be better understood from the brief description of figure, description, exemplary embodiments, example, figure, and claims that follow. Other desirable objectives

and advantages inherently achieved may occur or become apparent to those skilled in the art. The invention is defined by the appended claims.

BRIEF DESCRIPTION OF FIGURE

5 Figure 1 shows an optical micrograph of the silver nanowire product of Example 1.

DESCRIPTION

10 All publications, patents, and patent documents referred to in this document are incorporated by reference herein in their entirety, as though individually incorporated by reference.

 U.S. provisional application No. 61/488,837, filed May 23, 2011, entitled METAL ION CATALYSIS OF METAL ION REDUCTION, METHODS, COMPOSITIONS, AND ARTICLES, is hereby incorporated by
15 reference in its entirety.

Reducible Metal Ions and Metal Products

 Some embodiments provide methods comprising reducing at least one reducible metal ion to at least one metal. A reducible metal ion is a cation
20 that is capable of being reduced to a metal under some set of reaction conditions. In such methods, the at least one first reducible metal ion may, for example, comprise at least one coinage metal ion. A coinage metal ion is an ion of one of the coinage metals, which include copper, silver, and gold. Or such a reducible metal ion may, for example, comprise at least one ion of an IUPAC Group 11
25 element. IUPAC Group 11 elements are sometimes referred to as Group IB elements, based on historic nomenclature. An exemplary reducible metal ion is a silver cation. Such reducible metal ions may, in some cases, be provided as salts. For example, silver cations might, for example, be provided as silver nitrate.

 In such embodiments, the at least one metal is that metal to which
30 the at least one reducible metal ion is capable of being reduced. For example, silver would be the metal to which a silver cation would be capable of being reduced.

Nanostructures, Nanostructures, and Nanowires

In some embodiments, the metal product formed by such methods is a nanostructure, such as, for example, a one-dimensional nanostructure.

Nanostructures are structures having at least one “nanoscale” dimension less than
5 300 nm, and at least one other dimension being much larger than the nanoscale dimension, such as, for example, at least about 10 or at least about 100 or at least about 200 or at least about 1000 times larger. Examples of such nanostructures are nanorods, nanowires, nanotubes, nanopyramids, nanoprisms, nanoplates, and the like. “One-dimensional” nanostructures have one dimension that is much
10 larger than the other two dimensions, such as, for example, at least about 10 or at least about 100 or at least about 200 or at least about 1000 times larger.

Such one-dimensional nanostructures may, in some cases, comprise nanowires. Nanowires are one-dimensional nanostructures in which the two short dimensions (the thickness dimensions) are less than 300 nm, preferably
15 less than 100 nm, while the third dimension (the length dimension) is greater than 1 micron, preferably greater than 10 microns, and the aspect ratio (ratio of the length dimension to the larger of the two thickness dimensions) is greater than five. Nanowires are being employed as conductors in electronic devices or as elements in optical devices, among other possible uses. Silver nanowires are
20 preferred in some such applications.

Such methods may be used to prepare nanostructures other than nanowires, such as, for example, nanocubes, nanorods, nanopyramids, nanotubes, and the like. Nanowires and other nanostructure products may be incorporated into articles, such as, for example, electronic displays, touch screens, portable
25 telephones, cellular telephones, computer displays, laptop computers, tablet computers, point-of-purchase kiosks, music players, televisions, electronic games, electronic book readers, transparent electrodes, solar cells, light emitting diodes, other electronic devices, medical imaging devices, medical imaging media, and the like.

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Preparation Methods

A common method of preparing nanostructures, such as, for example, nanowires, is the “polyol” process. Such a process is described in, for example, *Angew. Chem. Int. Ed.* 2009, 48, 60, Y. Xia, Y. Xiong, B. Lim,

S. E. Skrabalak, which is hereby incorporated by reference in its entirety. Such processes typically reduce a metal cation, such as, for example, a silver cation, to the desired metal nanostructure product, such as, for example, a silver nanowire. Such a reduction may be carried out in a reaction mixture that may, for example, 5 comprise one or more polyols, such as, for example, ethylene glycol (EG), propylene glycol, butanediol, glycerol, sugars, carbohydrates, and the like; one or more protecting agents, such as, for example, polyvinylpyrrolidone (also known as polyvinylpyrrolidone or PVP), other polar polymers or copolymers, surfactants, acids, and the like; and one or more metal ions. These and other components may 10 be used in such reaction mixtures, as is known in the art. The reduction may, for example, be carried out at one or more temperatures from about 90 °C to about 190 °C.

Complexes Having Metal-Metal Bonds

15 The Applicant has recognized that molybdenum, as Mo^{2+} in a complex having a metal-metal quadruple bond can be used to prepare silver nanowires. An example of such a complex is $(\text{Mo}(\text{OCOCH}_3)_2)_2$. Such complexes are described in M.H. Chisholm, H.T. Chiu, J.C. Huffman, *Polyhedron*, **1984**, 3, 759, which is incorporated by reference in its entirety.

20 Other metal-metal bonded complexes that include metal-metal double bonds, triple bonds, or quadruple bonds, are thought to be useful in this application. Moreover, metal ions in other oxidation states, such as +3 or +4, such as, for example, Mo^{3+} or Mo^{4+} , are also expected to be useful in this application.

These methods are also believed to be applicable to reducible metal 25 cations other than silver cations, including, for example, reducible cations of other IUPAC Group 11 elements, reducible cations of other coinage metals, and the like. The method may also be used to prepare products other than nanowires, such as, for example, nanocubes, nanorods, nanopyramids, nanotubes, and the like. Such products may be incorporated into articles, such as, for example, transparent 30 electrodes, solar cells, light emitting diodes, other electronic devices, medical imaging devices, medical imaging media, and the like.

EXEMPLARY EMBODIMENTS

U.S. provisional application No. 61/488,837, filed May 23, 2011,
entitled METAL ION CATALYSIS OF METAL ION REDUCTION,
METHODS, COMPOSITIONS, AND ARTICLES, which is hereby incorporated

5 by reference in its entirety, disclosed the following 23 non-limiting exemplary
embodiments:

A. A method comprising:
providing a composition comprising:
at least one first compound comprising at least one first reducible metal

10 ion;

at least one second compound comprising at least a second metal atom and
at least a third metal atom, said at least one second metal atom being bonded to
the at least one third metal atom by at least one metal-to-metal bond; and
at least one solvent; and

15 reducing the at least one first reducible metal ion to at least one first metal.

B. The method of embodiment A, wherein the composition further comprises
at least one protecting agent.

C. The method of embodiment A, wherein the at least one first reducible
metal ion comprises at least one coinage metal ion.

20 D. The method of embodiment A, wherein the at least one first reducible
metal ion comprises at least one ion of an element from IUPAC Group 11.

E. The method of embodiment A, wherein the at least one first reducible
metal ion comprises at least one ion of silver.

F. The method of embodiment A, wherein the at least one first compound
25 comprises silver nitrate.

G. The method of embodiment A, wherein the at least one second metal atom
comprises molybdenum.

H. The method of embodiment G, wherein the at least one third metal atom
comprises molybdenum.

30 J. The method of embodiment A, wherein the at least one metal-to-metal
bond comprises one or more of at least one double bond, at least one triple bond,
or at least one quadruple bond.

K. The method of embodiment A, wherein the at least one metal-to-metal
bond comprises one or more of at least one triple bond or at least one quadruple

bond.

L. The method of embodiment A, wherein the at least one metal-to-metal bond comprises at least one quadruple bond.

5 M. The method of embodiment A, wherein the at least one second compound comprises at least one complex comprising the at least one second metal atom and the at least one third metal atom, said at least one complex further comprising at least one carbon atom bonded to the at least one second metal atom or the at least one third metal atom.

10 N. The method of embodiment A, wherein the at least one second compound comprises at least one complex comprising the at least one second metal atom and the at least one third metal atom, said at least one complex further comprising at least one first carbon atom bonded to the at least one second metal atom and at least one second carbon atom bonded to the at least one third metal atom.

15 P. The method of embodiment A, wherein the at least one second compound comprises $(\text{Mo}(\text{OCOCH}_3)_2)_2$.

Q. The method of embodiment A, wherein the at least one solvent comprises at least one polyol.

20 R. The method of embodiment A, wherein the at least one solvent comprises at least one of: ethylene glycol, propylene glycol, glycerol, one or more sugars, or one or more carbohydrates.

S. The method of embodiment A, wherein the composition has a ratio of the total moles of the at least one second metal atom to the total moles of the at least one first reducible metal ion from about 0.0001 to about 0.1.

25 T. The method of embodiment S, further wherein the composition has a ratio of the total moles of the at least one third metal atom to the total moles of the at least one first reducible metal ion from about 0.0001 to about 0.1.

U. The method of embodiment A, wherein the reduction is carried out at one or more temperatures from about 120 °C to about 190 °C.

30 V. The at least one first metal produced according to the method of embodiment A.

W. At least one article comprising the at least one first metal produced according to the method of embodiment A.

X. The at least one article of embodiment W, wherein the at least one first metal comprises one or more nanowires, nanocubes, nanorods, nanopyramids, or

nanotubes.

Y. The at least one article of embodiment W, wherein the at least one first metal comprises at least one object having an average diameter of between about 10 nm and about 500 nm.

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EXAMPLE

Example 1:

To a 500 mL reaction flask containing 280 mL ethylene glycol (EG), 2.0 g of a 5.1 mM solution of $(\text{Mo}(\text{OCOCH}_3)_2)_2$ in ethylene glycol (EG) and 10 0.15 g of a 1.3 mM solution of NaCl in EG were added. This solution was stripped of at least some dissolved gases by bubbling N_2 into the solution for at 2 hrs using a glass pipette at room temperature with mechanical stirring while at 100 rpm. (This operation will be referred to as “degassing” the solution in the sequel.) Solutions of polyvinylpyrrolidinone (PVP), 0.77 M in EG, and AgNO_3 , 15 0.25 M in EG, were also degassed with nitrogen for 60 min, then 20 mL syringes of each were prepared. The reaction mixture was heated to 145 °C under nitrogen and the AgNO_3 and PVP solutions were added at a constant rate over 25 minutes via a 12 gauge Teflon syringe needle. The reaction mixture was held at 145 °C for 90 minutes then allowed to cool to ambient temperature. Figure 1 shows an 20 optical micrograph of the product sample.

CLAIMS:

1. A method comprising:
providing a composition comprising:
at least one first compound comprising at least one first reducible
5 metal ion; and
at least one second compound comprising at least a second metal
atom and at least a third metal atom, said at least one second metal atom being
bonded to the at least one third metal atom by at least one metal-to-metal bond;
and
10 reducing the at least reducing the at least one first reducible metal
ion to at least one first metal nanowire.
2. The method according to claim 1, wherein the composition further
comprises at least one solvent.
- 15 3. The method of claim 1, wherein the at least one first reducible
metal ion comprises at least one coinage metal ion or at least one ion of an
element from IUPAC Group 11 or at least one silver ion..
- 20 4. The method of claim 1, wherein the at least one second metal atom
comprises molybdenum.
5. The method of claim 4, wherein the at least one third metal atom
comprises molybdenum.
- 25 6. The method of claim 1, wherein the at least one metal-to-metal
bond comprises one or more of at least one triple bond or at least one quadruple
bond.
- 30 7. The method of claim 1, wherein the at least one metal-to-metal
bond comprises at least one quadruple bond.
8. The method of claim 1, wherein the at least one second compound
comprises at least one complex comprising the at least one second metal atom and

the at least one third metal atom, said at least one complex further comprising at least one carbon atom bonded to the at least one second metal atom or the at least one third metal atom.

5 9. The method of claim 1, wherein the at least one second compound comprises $(\text{Mo}(\text{OCOCH}_3)_2)_2$.

10. The at least one first metal nanowire produced according to the method of claim 1.

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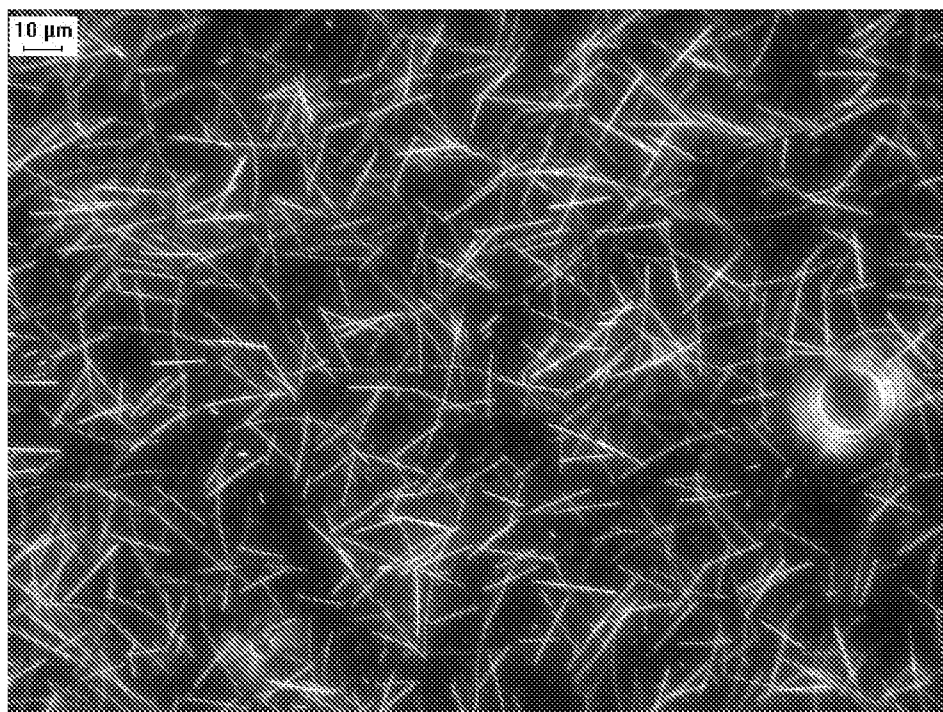


FIG. 1

INTERNATIONAL SEARCH REPORT

International application No

PCT/US2012/034876

A. CLASSIFICATION OF SUBJECT MATTER
INV. B22F9/24 B22F1/00
ADD.

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
B22F

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

EPO-Internal, WPI Data

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	JP 2011 032504 A (UNIV KYUSHU; DAINIPPON INK & CHEMICALS) 17 February 2011 (2011-02-17) example 1	1-3,6-8
A	----- WO 2010/136619 A2 (UNIV SEVILLA [ES]; CONSEJO SUPERIOR INVESTIGACION [ES]; ODRIOZOLA GORD) 2 December 2010 (2010-12-02) page 6, line 1 - line 21	4,5,9,10
X	----- EP 0 977 212 A2 (IBM [US]) 2 February 2000 (2000-02-02) paragraph [0086] - paragraph [0087]; example 1	1,2
A	----- JP 2009 155674 A (UNIV OSAKA) 16 July 2009 (2009-07-16) figures 1-15; examples 1-18	3-10
X	----- EP 0 977 212 A2 (IBM [US]) 2 February 2000 (2000-02-02) paragraph [0086] - paragraph [0087]; example 1	1,2
A	----- JP 2009 155674 A (UNIV OSAKA) 16 July 2009 (2009-07-16) figures 1-15; examples 1-18	3-10
X	----- JP 2009 155674 A (UNIV OSAKA) 16 July 2009 (2009-07-16) figures 1-15; examples 1-18	10
A	----- JP 2009 155674 A (UNIV OSAKA) 16 July 2009 (2009-07-16) figures 1-15; examples 1-18	1-9



Further documents are listed in the continuation of Box C.



See patent family annex.

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Date of the actual completion of the international search

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INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No

PCT/US2012/034876

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
JP 2011032504 A	17-02-2011	NONE	
WO 2010136619 A2	02-12-2010	ES 2352295 A1 WO 2010136619 A2	17-02-2011 02-12-2010
EP 0977212 A2	02-02-2000	AT 233938 T CN 1243775 A DE 69905663 D1 DE 69905663 T2 EP 0977212 A2 JP 3258295 B2 JP 2000054012 A KR 20000011546 A SG 83738 A1 US 6262129 B1	15-03-2003 09-02-2000 10-04-2003 20-11-2003 02-02-2000 18-02-2002 22-02-2000 25-02-2000 16-10-2001 17-07-2001
JP 2009155674 A	16-07-2009	NONE	