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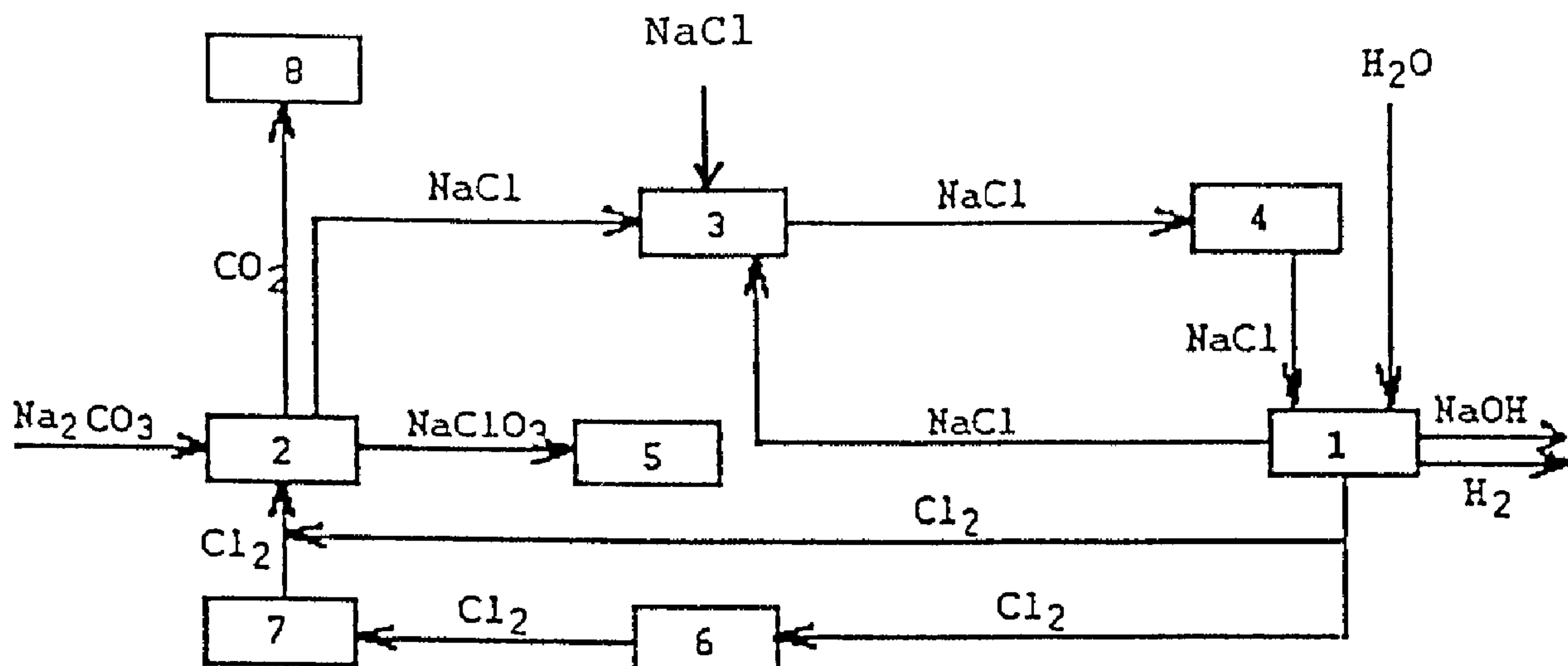
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(54) Titre : METHODE DE PREPARATION D'HYDROXYDES DE METAUX ALCALINS

(54) Title: METHOD FOR PRODUCING ALKALINE METAL HYDROXIDE



(57) Abrégé/Abstract:

The invention concerns a method for producing an alkali metal hydroxide by a chlor-alkali electrolysis (1), wherein an alkali metal chloride solution is electrolyzed in order to form alkali metal hydroxide, chlorine and hydrogen, wherein at least a part of the alkali metal chloride used in the electrolysis is prepared by neutralizing (2) an alkali metal carbonate with chlorine or hydrogen chloride.



Abstract

The invention concerns a method for producing an alkali metal hydroxide by a chlor-alkali electrolysis (1), wherein an alkali metal chloride solution is electrolyzed in order to form alkali metal hydroxide, chlorine and hydrogen, wherein at least a part of the alkali metal chloride used in the electrolysis is prepared by neutralizing (2) an alkali metal carbonate with chlorine or hydrogen chloride.

Method for producing alkaline metal hydroxide

The invention concerns a method for producing an alkali
5 metal hydroxide by chlor-alkali electrolysis wherein an
alkali metal chloride solution is electrolyzed to form an
alkali metal hydroxide, chlorine and hydrogen.

A problem of the chlor-alkali industries in many countries
10 is the imbalance between the demands for the alkali metal
hydroxide and chlorine which are the main products of the
electrolysis. It is expected that the imbalance will get
even worse in the future because of the environmental
pressure caused by the use of chlorine and its finished
15 products. The result of this is that alternative methods for
producing alkali metal hydroxides have to be found, because
they can't be produced in adequate quantities by the
conventional chlor-alkali electrolysis.

20 Other known methods for producing alkali metal hydroxides
include the following methods:
- chemical decarbonation of soda with either lime milk or
ammonia,
- electrolytical breaking down of an alkali metal sulphate
25 to an alkali metal hydroxide and sulphuric acid,
- decarbonation of soda with an acid solution of an alkali
metal sulphate and the electrolysis of so obtained sulphate
solution to an alkali metal hydroxide,
- catalytical breaking down of an alkali metal sulphide to
30 an alkali metal hydroxide and sulphur dioxide and
- electrolytical breaking down of an alkali metal chlorate
to an alkali metal hydroxide and chlorine dioxide.

The chlorate is generally produced by electrical breaking
35 down of alkali metal chlorides.

All of these known methods are unfavorable for the present producers of alkali metal hydroxide. The disadvantages of the most common alternative processes are the following: too much of sulphur compounds are produced as by-products in the process, having no use in a large scale; the raw material and energy costs of the process are too high compared to the market price of the product; big investments are required in the process and thus its capital costs are getting unfavorable in proportion to the market price of the product.

The intention of the invention is to provide a method for producing an alkali metal hydroxide, with which method the above disadvantages of the known methods are avoided and which can be carried out in chlor-alkali plants wherein additionally the ratio of the alkali metal hydroxide and the chlorine produced in the chlor-alkali electrolysis can be controlled to a desired level, and which method can additionally produce hydrogen and according to one embodiment also alkali metal chlorate, which products can be further utilized.

The principal characteristic features of the invention appear from the appended claims.

The invention resides in the realization that at least a part of the alkali metal chloride used in the chlor-alkali electrolysis can be produced by neutralizing or decarbonating an alkali metal carbonate either with chlorine or gaseous hydrogen chloride. By the decarbonation of an alkali metal carbonate with chlorine in an aqueous solution an alkali metal chlorate is at the same time formed.

By "alkali metal" it is in this connection primarily meant sodium and potassium, especially sodium.

When chlorine is used in the neutralization, the reaction products of the neutralization of an alkali metal carbonate in an aqueous solution are carbon dioxide, an alkali metal chloride and chlorate.

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When hydrogen chloride is used in the neutralization, the reaction products of the neutralization of an alkali metal carbonate are carbon dioxide, an alkali metal chloride and water.

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According to one preferred embodiment the alkali metal carbonate and chloride are fed in the neutralization stage to the reactor which contains water or a solution of an alkali metal chloride or an alkali metal chlorate and an alkali metal chloride. When neutralizing an alkali metal carbonate with chlorine an alkali metal chloride and chlorate are formed. The solution is getting supersaturated in respect of the chloride, and hence the alkali metal chloride can be separated from the solution as crystals.

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This alkali metal chloride is then used in a traditional way in the electrolysis in order to produce chlorine, alkali metal hydroxide and hydrogen. Preferably at least a part of the chlorine is circulated back to the above mentioned reactor. In order to avoid the supersaturation of the reactor solution in respect of chlorate a side stream is taken out from the reactor to the extended handling of chlorate.

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The crystalline alkali metal chloride obtained from the reactor is preferably dissolved to the liquid circulation of the chlor-alkali plant, i.e., to the diluted alkali metal chloride solution which is leaving the chlor-alkali electrolysis, wherein the concentration of the obtained solution is preferably near that of a saturated solution. This solution is after purification returned to the chlor-alkali electrolysis.

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According to a second preferred embodiment of the invention an alkali metal carbonate and chlorine are mixed in the neutralization stage to the diluted alkali metal chloride solution leaving the chlor-alkali electrolysis. An alkali metal chloride and chlorite are formed in the neutralization of the alkali metal carbonate by chlorine. The alkali metal chlorate is removed from the solution obtained or changed to the desired product by a known method, and the concentrated alkali metal chloride solution thus obtained is after purification returned to the chlor-alkali electrolysis to produce alkali metal hydroxide, chlorine and hydrogen. Preferably at least a part of the obtained chlorine is fed to the above mentioned neutralization stage.

According a third preferable embodiment of the invention the neutralization is carried out by adding the alkali metal carbonate to the alkali metal chloride solution which is leaving the electrolytic cells, to which alkali metal chloride solution gaseous hydrogen chloride has been absorbed, and the alkali metal chloride solution thus obtained is then fed after purification to the electrolytic cells.

According a fourth preferable embodiment of the invention the neutralization is carried out by adding the alkali metal carbonate to the alkali metal chloride solution which is leaving the electrolytic cells and by neutralizing the carbonate thereafter with hydrogen chloride gas, and the alkali metal chloride solution thus obtained is then after purification fed to the electrolytic cells.

According a fifth preferable embodiment of the invention the neutralization is carried out in a closed circulation containing initially either water or alkali metal chloride solution. To this circulation is then added alkali metal carbonate and either hydrogen chloride gas or hydrochloric acid. When the alkali metal carbonate is decarbonated the

solution is getting supersaturated in respect of alkali metal chloride which can then be separated from the solution as crystals. These crystals are then led to the solution circulation of the chlor-alkali electrolysis in order to produce chlorine, an alkali metal hydroxide and hydrogen.

Consequently in the neutralization is produced the alkali metal chloride solution of a conventional chlor-alkali plant, which alkali metal chloride solution is purified by known processes depending on the electrolytic cell method in question.

The alkali metal chloride solution to be fed to the chlor-alkali electrolysis is preferably nearly saturated.

The pH value of the solution in the neutralization is preferably over 3 and especially between 3 and 11.

The neutralization can be carried out either continuously or batchwise at a broad temperature range which is preferably between 20°C and 100°C. When chlorine is used to the neutralization the alkali metal chlorate and chloride concentrations of the solution to be taken out is influenced by the temperature. The neutralization can be carried out in one or several stages.

According to the method of the invention the alkali metal chloride is electrolyzed in a known manner in the electrolytic cells in order to produce the alkali metal hydroxide. The electrolysis can be carried out in mercury cathode, diaphragm or membrane cells. The quality of the alkali metal hydroxide solution obtained after the electrolysis is equivalent to the quality obtained by the conventional chloride method.

The chlorine gas used in the neutralization of the alkali metal carbonate is preferably produced by the electrolysis.

Chlorine is decarbonating a natural carbonate in the solution forming again alkali metal chloride, alkali metal chlorate and carbon dioxide. The carbon dioxide can be purified further and liquefied in a known manner for further use. The alkali metal chlorate can be fed to the preparation of alkali metal chlorate, or to a process where alkali metal chlorate is spent, or it can be changed to a chloride solution with hydrochloric acid, or it can be destroyed, or it can be refined further by some other desired way.

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Hydrogen chloride which is used in the neutralization of alkali metal carbonate is preferably produced from the chlorine and hydrogen gases produced in the synthesis apparatus, to which also a known excess of hydrogen gas needed for the hydrogen chloride synthesis is led. After the synthesis the obtained hydrogen-containing hydrogen chloride gas can be absorbed in a known manner either in water or a diluted hydrochloric acid solution. By heating a concentrated hydrochloric acid solution a pure hydrogen chloride gas is provided which is absorbed to the diluted alkali chloride solution leaving the electrolysis cells. The obtained acidic chloride solution is decarbonating the natural carbonate, forming again alkali metal chloride and carbon dioxide. The carbon dioxide can be purified further and liquefied in a known manner for further use.

25

A remarkable advantage of the present invention is that already existing but underutilized chlor-alkali plants can be utilized in the process. By controlling the amount of the alkali metal carbonate it is according to the invention possible to influence on how much of the chlorine produced by the electrolysis is remaining as saleable product of the process. One of the essential characteristics of the invention is particularly that the ratio of chlorine to be saled to the production of the alkali metal hydroxide can be regulated steplessly from about 0% to nearly 100% by varying the amount of salt obtained from the alkali metal carbonate

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and to be fed to the process. The rest of the alkali metal chloride needed in the electrolysis is provided by feeding new alkali metal chloride to the solution circulation.

- 5 Another important advantage of the invention is that in the chlorination of the carbonate preferably chlorate is formed which can be utilized easily and economically. By adjusting the temperature in the reactor the chlorate and chloride contents of the obtained solution can be influenced and
10 hence obtain a composition of the solution which is best suited for the intended use.

A third important advantage of the invention is that the hydrogen formed in the chlor-alkali electrolysis can
15 advantageously be utilized either in other chemical processes or pro-environmentally as a fuel in the energy production.

The neutralization degree of the alkali metal carbonate can
20 be adjusted by adjusting the amount of hydrogen or hydrogen chloride and hence to provide suitable conditions in respect of the purification and electrolysis of the alkali metal chloride solution.

- 25 Among the most remarkable advantages in respect of the existing chlor-alkali plants the following can be mentioned: In the method a cheap, even unrefined alkali metal carbonate can be used which is suitable for electrolytic use. Noxious sulphur compounds are not produced in the method as by-
30 products. The existing capacity of an electrolytic cell can be utilized to its full value. The investments which the method needs in existing plants are very small compared to alternative processes with the same production capacity of alkali metal hydroxide. The method is very flexible when the
35 the production demand for chlorine and lye is changing. The production costs of the alkali metal hydroxide realized by means of the present invention are very low compared to the

alternative processes. When chlorine is used to the neutralization the method is producing chlorate and hydrogen as by-products which can be utilized economically, and hence the method is economically more uniform.

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The invention is now illustrated in more detail in connection of the appended Figures wherein:

10 Fig. 1 is a block diagram showing the principles of the production of an alkali metal hydroxide according to the invention, as adapted for the production of lye,

15 Fig. 2 is a block diagram showing the principles of a second production of an alkali metal hydroxide according to the invention, as adapted for the production of lye,

20 Fig. 3 is a block diagram showing the principles of a third production of an alkali metal hydroxide according to the invention, as adapted for the production of lye,

Fig. 4 is a block diagram showing the principles of a fourth production of an alkali metal hydroxide according to the invention, as adapted for the production of lye.

25 In Fig. 1 the electrolysis is marked with the reference numeral 1. To the electrolysis is fed a sodium chloride solution through the purification 4. From the electrolysis 1 the desired sodium hydroxide solution and hydrogen are obtained which can be treated further in a method which is
30 known per se. The chlorine gas produced by the electrolysis 1 is transported to the reactor 2, where the chlorine gas can also be taken through liquefaction 6 and vaporization 7. In reactor 2 the sodium chloride crystal formed as the reaction product of chlorine and soda is separated from the
35 reactor solution and is at saturating station 3 dissolved to the diluted salt solution returning from the electrolysis 2 and where also additional sodium chloride is added as

needed. The saturated salt solution is led from the saturating station 3 to the purification 4, and from there further to the electrolysis 1. In reactor 2 there are also carbon dioxide CO_2 and sodium chlorate, which are led to the extended treatments 5 and 8. The extended treatment 5 can be a separating of sodium chlorate from the solution or its utilization as such in other processes.

In Fig. 2 a sodium chloride solution is fed to the electrolysis 1 through the purification 4. From the electrolysis 1 the desired sodium hydroxide solution and hydrogen are obtained, which can be further treated by a method known per se. The sodium carbonate, and chlorine produced by the electrolysis 1, are fed directly to the solution circulation of the chlor-alkali electrolysis in reactor 2. The chlorine gas can also be led to the reactor 2 through liquefaction 6 and vaporization 7. In reactor 2 the chlorate formed to the solution circulation is transformed by a known method to a desired compound at point 9, from where the concentrated salt solution is led to the purification 4 and from there further to the electrolysis 1. When chlorine is desired as a sales product additional sodium chloride is added to the solution circulation e.g. at point 2.

In Fig. 3, the electrolysis is indicated by the reference numeral 1. To the electrolysis a sodium chloride solution is fed which has been obtained from the sodium chloride purification 4. From the electrolysis 1 a desired sodium hydroxide solution is obtained which can be treated further by a method known per se. The chlorine and hydrogen gases produced by the electrolysis 1 are taken to the hydrochloric acid synthesis 10, from where the hydrochloric acid is led to the separation of hydrogen chloride gas 11. The pure hydrogen chloride gas obtained is absorbed in the absorption of hydrogen chloride 12 to the diluted sodium chloride solution which is leaving the electrolysis. The acidic sodium

chloride solution obtained is led to the neutralization 2, where even soda is fed. In the neutralization 2 the acidic chloride solution is decarbonating the soda wherein sodium chloride is formed which is led to the purification of the sodium chloride 4, and carbon dioxide, which is fed to the purification of carbon dioxide 8.

In Fig. 4, the electrolysis is indicated by the reference numeral 1. In it a sodium chloride solution is fed from the saturation 3, to which, when producing chlorine to be sold, external sodium chloride is brought. From electrolysis 1 a desired sodium hydroxide is obtained which can be treated further by a method known per se. The chlorine and hydrogen gases produced by electrolysis 1 are taken to the hydrochloric acid synthesis 10. From here the hydrochloric acid can be led as such or through the vaporization 11 to the reactor 12. In reactor 12 the hydrochloric acid is decarbonating soda forming sodium chloride, carbon dioxide and water. The solution will get supersaturated in respect of chloride, and sodium chloride can be separated from the solution as crystals. These crystals are then led to the saturation 3. From the reactor 12 the salt solution is returned to the dissolution of soda 2, from where it is led further through the purification 4 to the reactor 12. The carbon dioxide formed is purified in point 8 and is led to further treatment.

The workability of the process has been tested with several laboratory scale tests. In these the goal has been to find out the influence of temperature, pH and different concentrations on the progress of the reaction. According to the results, chlorine is produced five or over five times as much as chlorate. The end-pH is, depending on the chlorination degree, over 3. According to the chlorate/chloride-solubility curve the temperature and concentrations have a correlation which dictates the composition of the solution to be removed from the reactor.

In the following the invention is explained with examples.

Example 1

250 ml of a solution was taken, containing

5	NaCl	170g/l
	NaClO ₃	400g/l
	Na ₂ CO ₃	50g/l
	Temperature	< 30°C

10 Chlorine gas was fed through this solution until all sodium carbonate had reacted with chlorine. The residual chlorine was destroyed from the solution thus obtained by mixing. The chlorinated solution and the crystals formed were analyzed. As a result a solution was obtained containing

15	NaCl	185.2g/l
	NaClO ₃	406.2g/l

9.3g of a crystalline salt was obtained, containing 0.42g NaClO₃ and 8.92g NaCl. The final pH was 4.7

Example 2

Into 1 litre of water 100g Na₂CO₃ was dissolved. The pH of the obtained solution was 10.2. This solution was
25 chlorinated until Na₂CO₃ had been spent. As a result a solution was obtained containing 83.9g NaCl, 25.5g NaClO₃ and 5.5g active chlorine. Final pH was 6.2

30 Example 3

One thousand millilitres of a sodium chloride solution was taken having a temperature of 65°C and concentration 253g NaCl/l. To the solution 30 g hydrogen chloride gas was slowly absorbed which had been produced by heating a 33%
35 hydrochloric acid. To the acidic salt solution 45g of technical sodium carbonate having a concentration of 99.3% Na₂CO₃ was slowly added. After cessation of the carbon

dioxide evolution a salt concentration 298g/l was analyzed. The final temperature of the solution was 55°C and final pH 5. According to the analysis 49g sodium chloride had formed, well in accordance with the theoretical calculations. The
5 solution obtained has such a quality that it can be used as the feed solution of electrolysis.

CLAIMS

1 1. A method for producing alkali metal hydroxide by chlor-alkali
2 electrolysis which comprises the steps:
3 (a) providing a concentrated alkali metal chloride feed solution,
4 (b) electrolyzing said feed solution to form an alkali metal hydroxide
5 solution, chlorine, hydrogen and a diluted alkali metal chloride solution,
6 (c) recovering said alkali metal hydroxide solution,
7 (d) adding alkali metal chloride to said diluted alkali metal chloride
8 solution to provide a concentrated alkali metal chloride feed solution, and
9 (e) returning said feed solution obtained in step (d) to the chlor-alkali
10 electrolysis,
11 characterized in that at least a part of the alkali metal chloride added to said
12 diluted alkali metal chloride solution in step (d) is prepared by neutralizing an
13 aqueous solution of natural alkali metal carbonate with chlorine or hydrogen
14 chloride arising from respectively the chlorine and chlorine and hydrogen
15 products of the electrolysis step (b).

1 2. Method according to claim 1, in which the neutralization is carried
2 out by adding the alkali metal carbonate to water or to a solution containing
3 either an alkali metal chloride or an alkali metal chloride and alkali metal
4 chlorate, and thereafter feeding chlorine produced in step (b) to the solution to
5 precipitate alkali metal chloride crystals, whereafter said crystals are separated
6 from the solution and added to the diluted alkali metal chloride solution from
7 the chlor-alkali electrolysis to provide a concentrated solution which is purified
8 and returned to the chlor-alkali electrolysis.

1 3. Method according to claim 2, in which a part of the chlorate-
2 chloride solution formed in the neutralization is used again in the neutralization.

1 4. Method according to claim 1 or 2, in which the neutralization of
2 the alkali metal carbonate is carried out with chlorine produced in step (b) to
3 form a solution containing alkali metal chloride and alkali metal chlorate, and
4 in which at least a part of the chlorate-chloride solution formed in the
5 neutralization is supplied to a process for preparing alkali metal chlorate or to

6 a process where alkali metal chlorate is used or to a process where the alkali
7 metal chlorate is transformed to a chloride or to a process where the alkali
8 metal chlorate is destroyed.

1 5. Method according to claim 1, in which the neutralization is carried
2 out by adding the alkali metal carbonate to the diluted alkali metal chloride
3 solution from the chlor-alkali electrolysis, and thereafter feeding chlorine
4 produced in step (b) to the solution to provide a concentrated alkali metal
5 chloride solution containing alkali metal chlorate, whereafter the alkali metal
6 chlorate is removed from the concentrated solution and the obtained
7 concentrated solution is returned to the chlor-alkali electrolysis.

1 6. Method according to claim 5, in which the alkali metal chlorate in
2 the concentrated solution to be returned to the chlor-alkali electrolysis is
3 transformed to a desired compound.

1 7. Method according to claim 1, in which the neutralization is carried
2 out by absorbing hydrogen chloride gas produced from the hydrogen and
3 chlorine produced in step (b), into the diluted alkali metal chloride solution from
4 the chlor-alkali electrolysis and thereafter adding the alkali metal carbonate to
5 the solution to provide a concentrated alkali metal chloride solution, whereafter
6 said concentrated solution is purified and returned to the chlor-alkali
7 electrolysis.

1 8. Method according to claim 1, in which the neutralization is carried
2 out by adding the alkali metal carbonate to the diluted alkali metal chloride
3 solution from the chlor-alkali electrolysis and thereafter feeding hydrogen
4 chloride gas produced from the hydrogen and chlorine produced in step (b), to
5 the solution to provide a concentrated alkali metal chloride solution, whereafter
6 said concentrated solution is purified and returned to the chlor-alkali
7 electrolysis.

1 9. Method according to claim 1, in which the neutralization is carried
2 out in a closed solution circulation to which the alkali metal carbonate and
3 hydrochloric acid or hydrogen chloride gas, both produced from the hydrogen

4 and chlorine produced in step (b), is fed, and from which alkali metal chloride
5 crystals formed are separated and added to the diluted alkali metal chloride
6 solution from the chlor-alkali electrolysis to provide a concentrated alkali metal
7 chloride solution which is returned to the chlor-alkali electrolysis.

1 10. Method according to any one of claims 1-9, in which the
2 concentration of the alkali metal chloride solution to be fed to the chlor-alkali
3 electrolysis is about equal to that of the concentration of a saturated chloride.

1 11. Method according to any one of claims 1-10, in which pH of the
2 solution in the neutralization is over 3.

1 12. Method according to any one of claims 1-11, in which the
2 neutralization is carried out in one or several stages continuously or batchwise.

1 13. Method according to any one of claims 1-12, in which the
2 neutralization temperature is between 20°C and 100°C.

1 14. Method according to any one of claims 1-13, in which the
2 electrolysis of the alkali metal chloride solution is carried out in a mercury
3 cathode cell, a diaphragm cell or a membrane cell.

1 15. Method according to claim 1, in which chlorine gas is produced
2 in step (b), which chlorine gas is purified by liquefying and evaporating the
3 chlorine gas prior to feeding said chlorine gas to the neutralization.

1 16. Method according to claim 1, in which chlorine gas is produced
2 in step (b), which chlorine gas is fed directly without any purification to the
3 neutralization of the alkali metal carbonate.

1 17. Method according to claim 1, in which chlorine and hydrogen
2 gases produced in step (b) are reacted to produce hydrogen chloride gas, which
3 hydrogen chloride gas is purified by liquefying and evaporating the hydrogen
4 chloride gas.

1 18. Method according to claim 1, in which chlorine and hydrogen
2 gases produced in step (b) are reacted to produce hydrogen chloride gas, which
3 hydrogen chloride gas is purified by absorbing the hydrogen chloride gas into
1 water or into a hydrochloric acid solution to provide a hydrochloric acid
2 solution, and distilling said acid solution to yield pure hydrogen chloride gas.

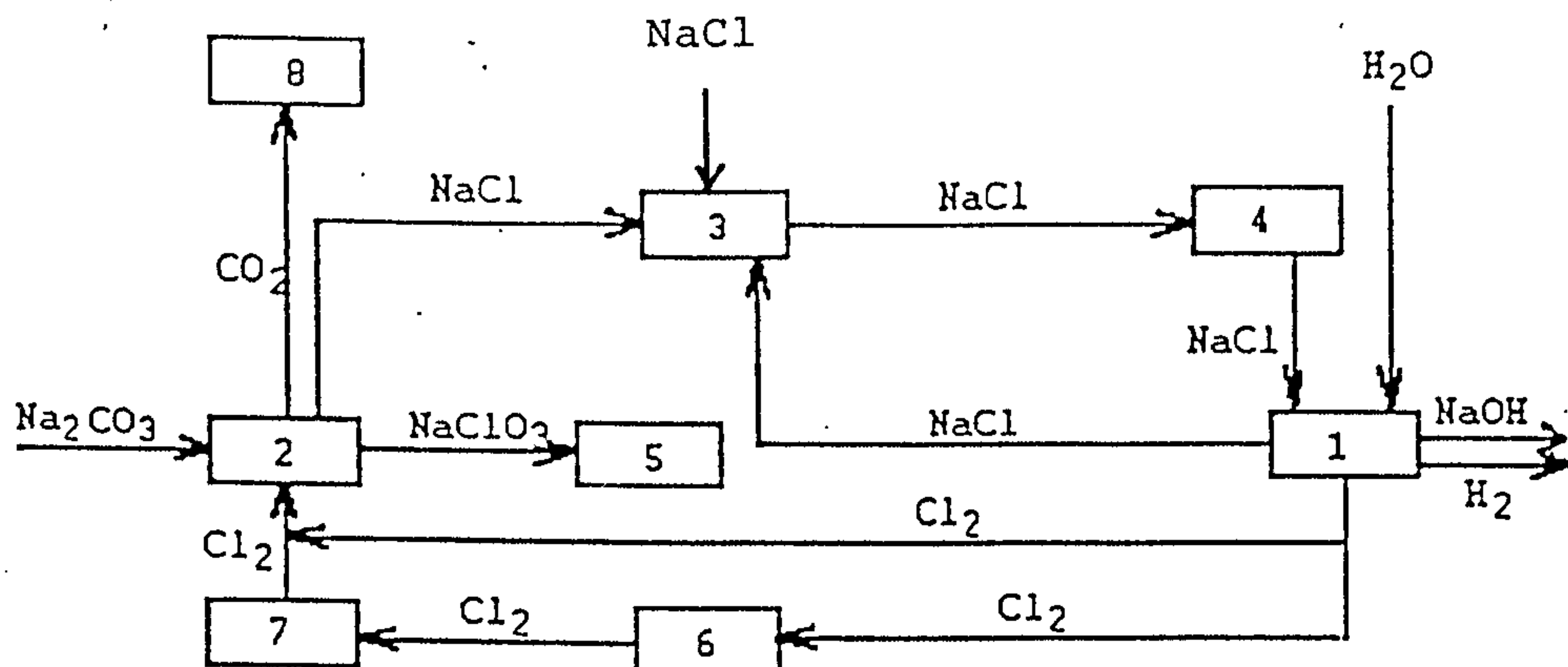


Fig. 1

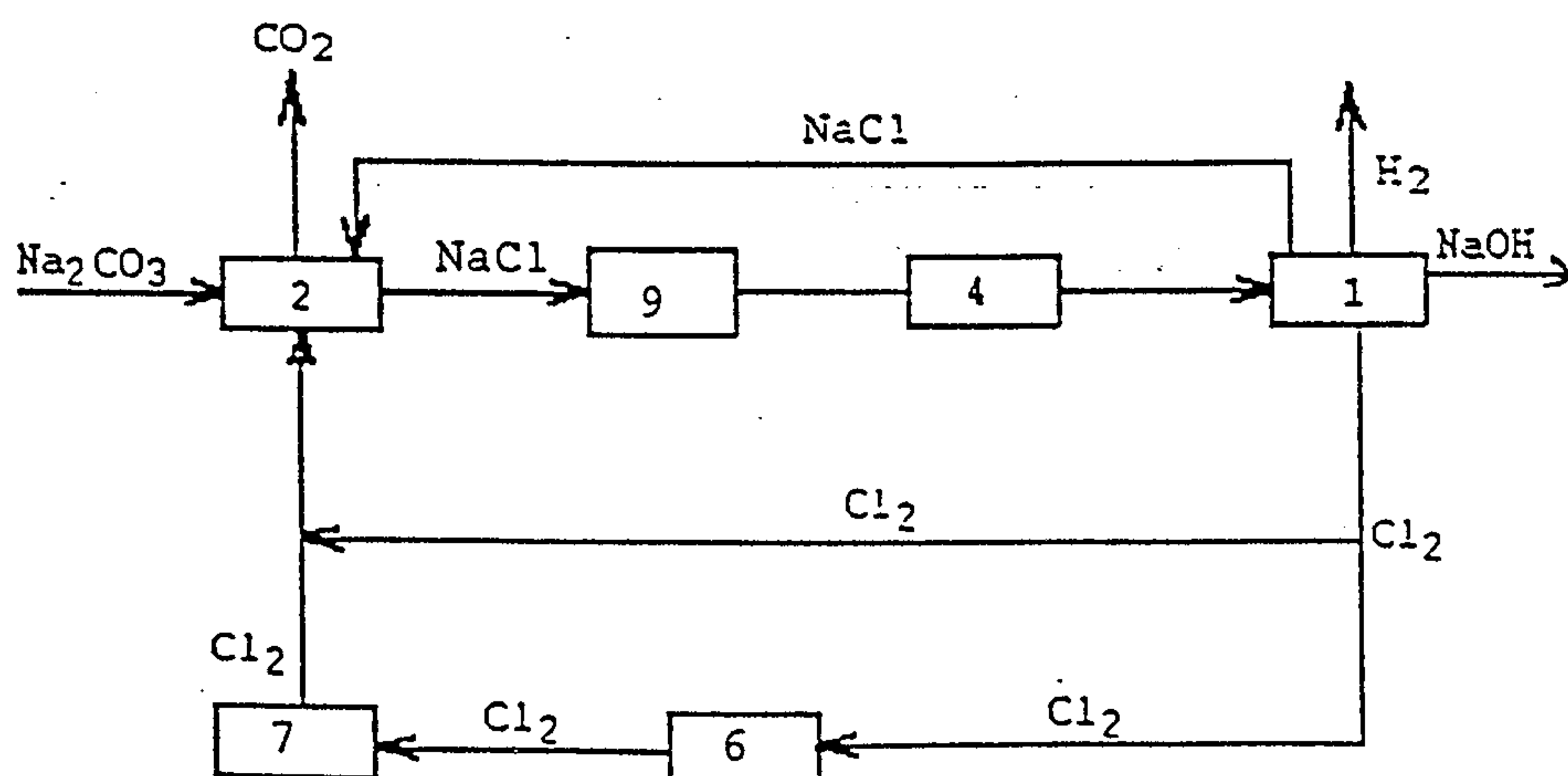


Fig. 2

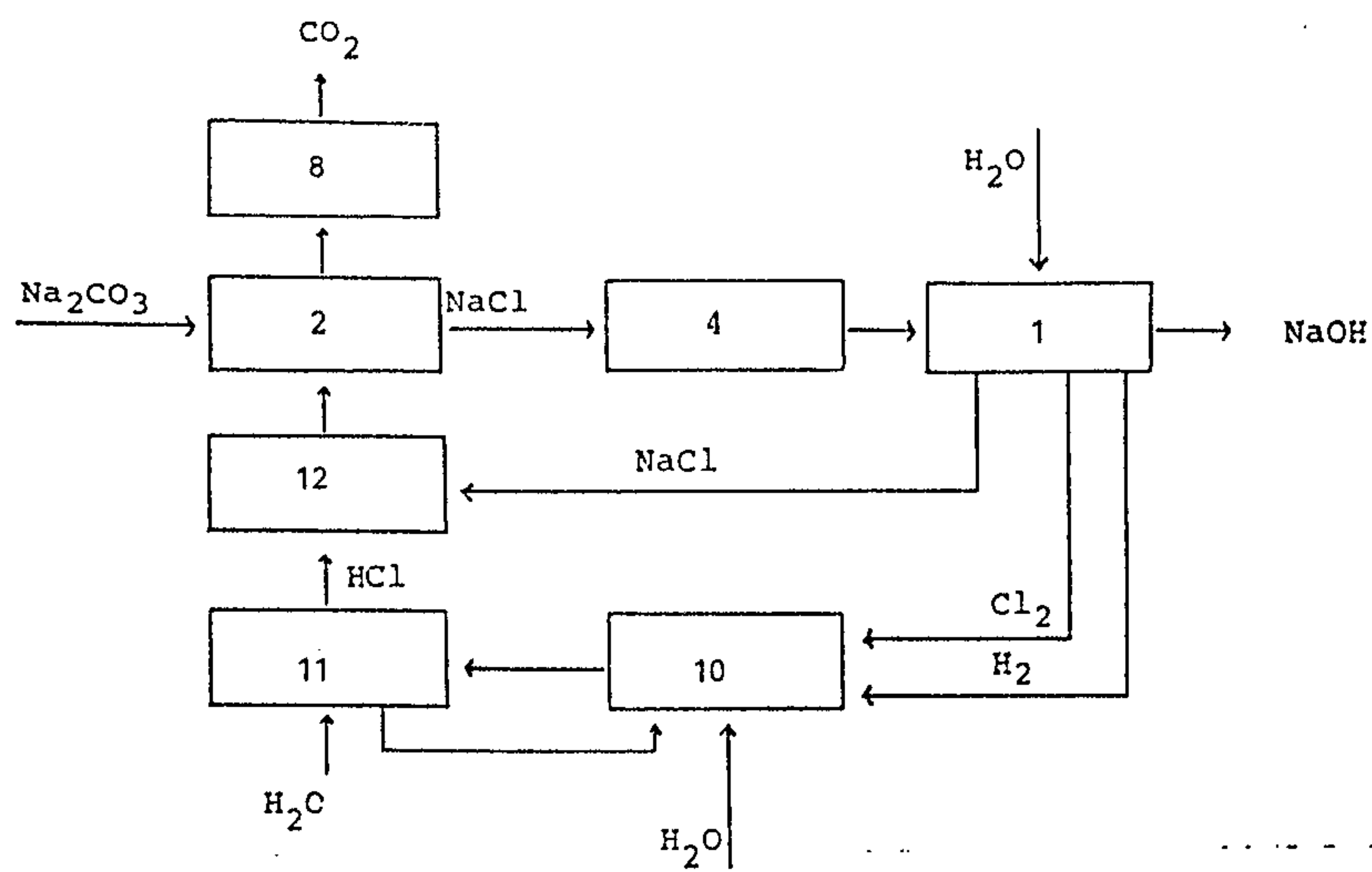


Fig. 3

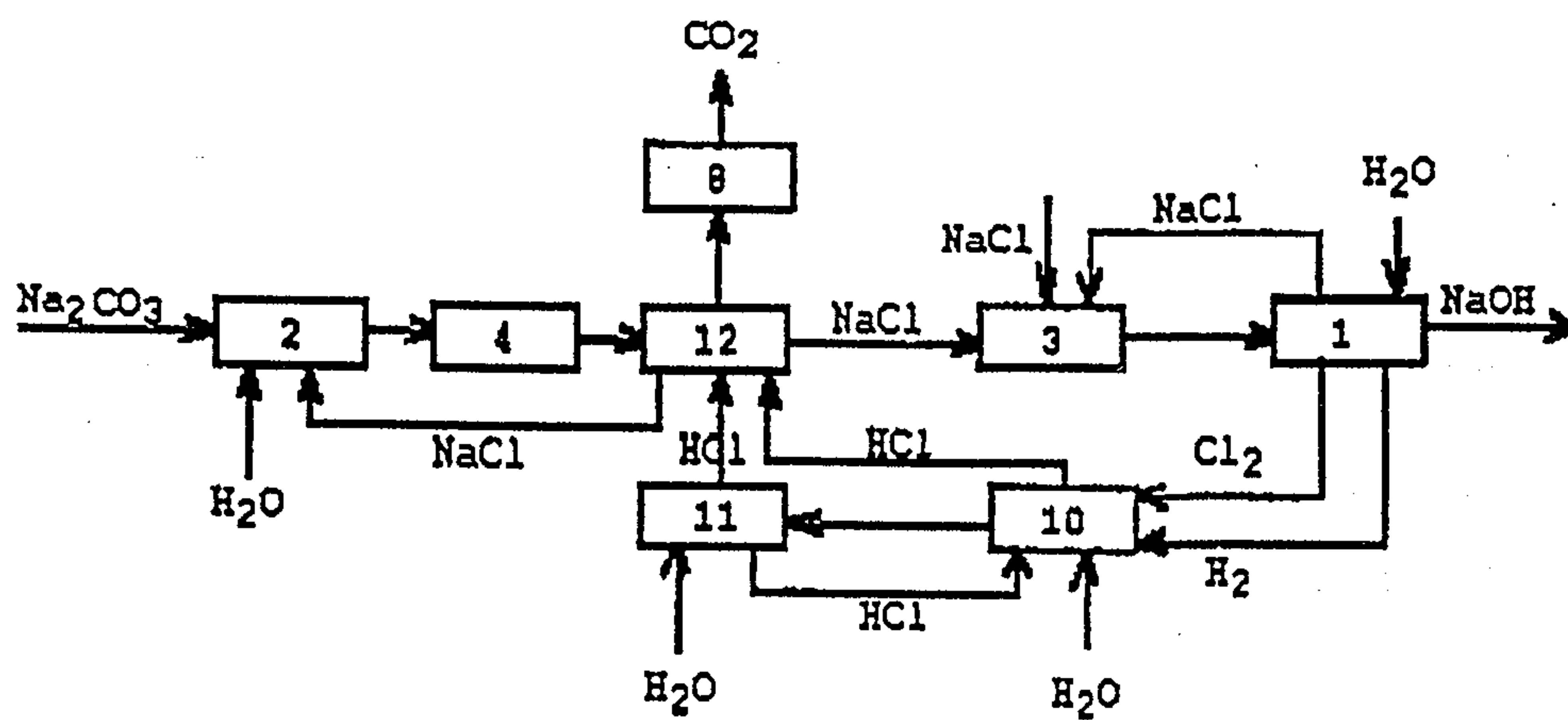


Fig. 4

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