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(54) Title: N-VINYLPYRROLIDONE-BASED CATIONIC COPOLYMER FOR SEPARATING AN OIL-IN-WATER
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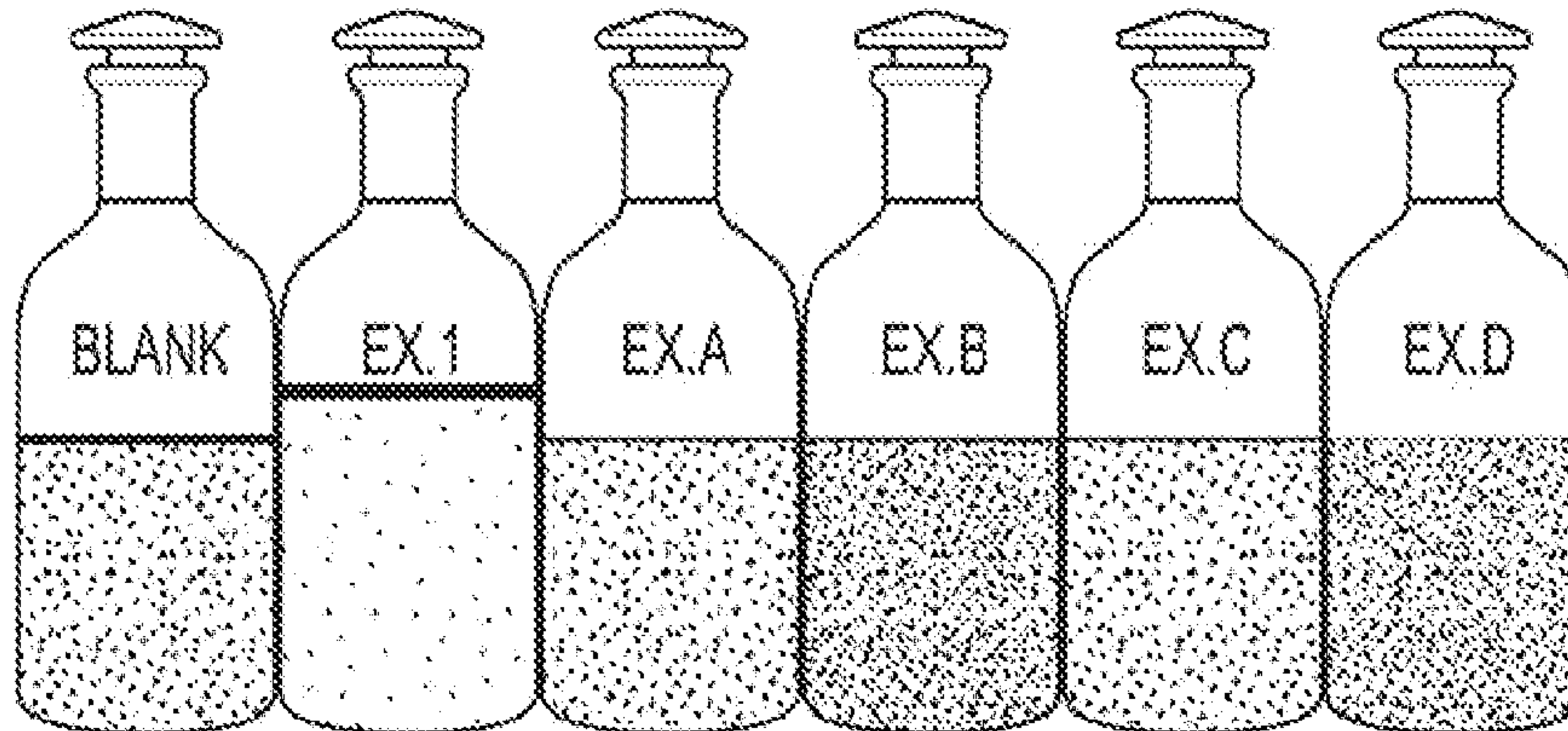


FIG. 1C

(57) **Abrégé/Abstract:**

Embodiments of the present disclosure include a method of separating an oil-in-water emulsion formed during crude oil production into a water phase and an oil phase that includes adding 1 part-per-million (ppm) to 10000 ppm of an N-vinylpyrrolidone based cationic copolymer to the oil-in-water emulsion, based on the total volume of the oil-in-water emulsion, to form a water phase and an oil phase, and separating the water phase from the oil phase.



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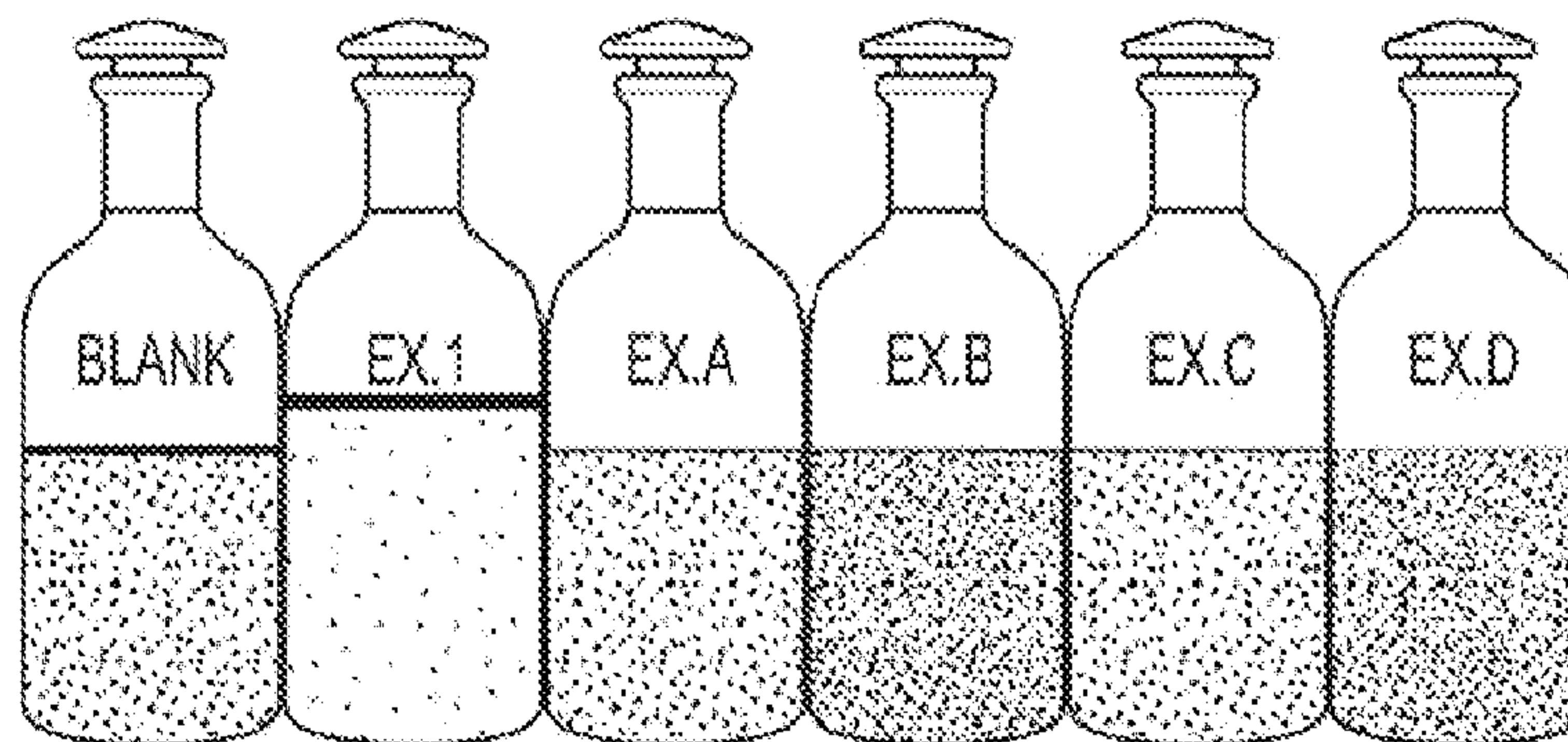


FIG. 1C

(57) Abstract: Embodiments of the present disclosure include a method of separating an oil-in-water emulsion formed during crude oil production into a water phase and an oil phase that includes adding 1 part-per-million (ppm) to 10000 ppm of an N-vinylpyrrolidone based cationic copolymer to the oil-in-water emulsion, based on the total volume of the oil-in-water emulsion, to form a water phase and an oil phase, and separating the water phase from the oil phase.

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N-VINYLPYRROLIDONE-BASED CATIONIC COPOLYMER FOR SEPARATING AN OIL-IN-WATER EMULSION

Field of the Disclosure

5 The present disclosure relates to a method of separating an oil-in-water emulsion and more particularly to a method of separating an oil-in-water emulsion formed during crude oil production.

Background

10 Chemical aids used to remove dispersed oil and/or solids from water are commonly referred to as water clarifiers, reverse emulsion breakers, deoilers, coagulants, flocculants and/or coalescence aids. In the oil and gas industry, after the initial separation of the bulk produced fluids (e.g., crude oil), the produced water still contains finely dispersed solids and oil. These oil and solids particles are well stabilized and are difficult to separate by means of physical settling
15 alone. Often, such produced water cannot be reused nor disposed of as is and it is therefore necessary to find appropriate solutions to do so. Regulations around the world generally limit the oil and grease content in produced water to a maximum of 15 parts-per-million (ppm) to 50 ppm for discharge into the environment (Arnold, K.; Stewart, M. Surface Production Operations; 3rd ed.; Elsevier/Gulf: Boston, 2008, 483). The water is thus treated to meet regulatory,
20 environmental, and operational goals.

 A range of synthetic water soluble cationic polymers are known to separate oil/solid particles from produced water. Poly(diallyl dimethylammonium chloride), copolymers of acrylamide or alkyl acrylates with various cationic co-monomers are known as water clarifiers. Chemical treatment of the produced water involves the addition of a few ppm levels of inorganic
25 salts and/or organic polymers to facilitate the separation through coagulation and floc formation. Organic polymers (ionic or neutral) can be more effective than the inorganic salts and results in water with minimum oil/solid residues.

Summary

30 The present disclosure provides a method of separating an oil-in-water emulsion formed during crude oil production into a water phase and an oil phase using a cationic copolymer. As

discussed herein, the oil-in-water emulsion includes droplets of oil, such as crude oil. Specifically, embodiments of the present disclosure include a method of separating the oil-in-water emulsion formed during crude oil production into a water phase and an oil phase that includes adding 1 part-per-million (ppm) to 10000 ppm of an N-vinylpyrrolidone-based cationic copolymer to the oil-in-water emulsion, based on the total volume of the oil-in-water emulsion, to form a water phase and an oil phase, and separating the water phase from the oil phase.

Brief Description of Drawings

Fig. 1A is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 1, treated with 0 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 6.9).

Fig. 1B is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 1, treated with 100 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 6.9).

Fig. 1C is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 1, treated with 300 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 6.9).

Fig. 2A is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 2, treated with 0 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 7.7).

Fig. 2B is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 2, treated with 50 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 7.7).

Fig. 2C is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 2, treated with 100 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 7.7).

Fig. 3A is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 3, treated with 0 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 8.9).

Fig. 3B is a photographic image of the synthetic produced oil-in-water emulsion for Bottle Test Procedure 3, treated with 50 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 8.9).

Fig. 3C is a photographic image of the synthetic produced oil-in-water emulsion for
5 Bottle Test Procedure 3, treated with 100 ppm of cationic polymer Example 1 and Comparative Examples A-D (pH about 8.9).

Detailed Description

The present disclosure provides a method of separating an oil-in-water emulsion,
10 commonly referred to in the oilfield industry as a reverse emulsion, using a cationic copolymer, where the cationic copolymer helps to break the emulsion into a water phase and an oil phase. The method of the present disclosure is particularly useful in crude oil production and processing systems including refinery water treatment and even petrochemical plants.

As discussed herein, an oil-in-water emulsion can include droplets of oil, such as crude
15 oil. The oil-in-water emulsion can also include, besides oil droplets, solid particles, such as clays, silts, sand, corrosion by-products, and scale, among other solid particles which can be present in the emulsion. As discussed herein, separating the oil-in-water emulsion is a process in which the emulsion is broken into its constituents of a water phase and an oil phase. As used herein, the "water" of the oil-in-water emulsion and/or the water phase can include, for example,
20 a brine, a connate water, surface water, steam condensate, carbonated water, sea water and a combination thereof. For brevity, the word "water" is used herein, where it is understood that "brine," "connate water," "surface water," "steam condensate," "carbonated water," and/or "sea water" can be used interchangeably where needed.

The oil-in-water emulsion can be produced in extracting crude oil (a naturally occurring
25 flammable liquid found in geological formations beneath the earth's surface, which consists of a complex mixture of hydrocarbons of various molecular weights). An oil-in-water emulsion can be formed in almost every phase of crude oil production and processing. As used herein, "oil-in-water" emulsions can include oil-in-water (e.g., a reverse emulsion) and multiple or complex emulsions, as are known, where the oil forms the dispersed phase and the water forms the
30 continuous phase. Droplets or particles of the oil-in-water emulsion can, but need not, vary in size from 1 micrometer to 1000 micrometer. Droplets or particles of less than 1 micrometer

and/or greater than 1000 micrometer are also possible. As used herein, resolution, separation or reverse demulsification means the breaking of an oil-in-water emulsion into an oil phase and a water phase.

The cationic copolymer of the present disclosure is an N-vinylpyrrolidone-based cationic copolymer. For the various embodiments, the N-vinylpyrrolidone-based cationic copolymer is formed with monomers of N-vinyl-2-pyrrolidone and a quaternized dimethylaminoethyl methacrylate. The quaternized dimethylaminoethyl methacrylate can consist of 2-ethyldimethylaminoethyl methacrylate ethyl sulfate to provide a poly[(2-ethyldimethylaminoethyl methacrylate ethyl sulfate)-co-N-vinyl-2-pyrrolidone]. For the various embodiments, the N-vinylpyrrolidone-based cationic copolymer can be supplemented with low levels of additives without impairing separation performance. The additive can consist of, but is not limited to, biocidal preservatives and can be included at levels up to 1 wt.% relative to the total weight of the cationic copolymer/additive mixture. In one embodiment, the N-vinylpyrrolidone-based cationic copolymer is a copolymer of N-vinyl-2-pyrrolidone (67 weight percent (wt.%) based on the total weight of the copolymer) and 2-ethyldimethylaminoethyl methacrylate ethyl sulfate (33wt.%).

One such N-vinylpyrrolidone-based cationic copolymer is poly[(2-ethyldimethylaminoethyl methacrylate ethyl sulfate)-co-N-vinyl-2-pyrrolidone] formed as a reaction product of about 67 wt.% of N-vinyl-2-pyrrolidone and about 33 wt.% of 2-ethyldimethylaminoethyl methacrylate ethyl sulfate. One such N-vinylpyrrolidone-based cationic copolymer is available from BASF under the trade name LUVIQUAT® PQ 11.

The cationic copolymer of the present disclosure can be prepared by solution, emulsion, or dispersion polymerization techniques. The weight average molecular weight of the N-vinylpyrrolidone-based cationic copolymer can range from 5000 grams/mole (g/mol) to 5000000 g/mol and preferably ranges from 20000 g/mol to 2000000 g/mol. In one embodiment, the N-vinylpyrrolidone-based cationic copolymer of the present disclosure has a weight average molecular weight of about 1000000 g/mol. Weight average molecular weight can be measured by gel permeation chromatography or small-angle dynamic light scattering.

Using the cationic copolymer of the present disclosure in separating an oil-in-water emulsion formed during crude oil production into a water phase and an oil phase may be carried out in a conventional manner. For example, separating the oil-in-water emulsion into an oil

phase and a water phase and then recovering the oil phase and water phase may be carried out by treating the oil-in-water emulsion with a separating amount of the cationic copolymer of the present disclosure. Examples of separating the oil-in-water emulsion formed during crude oil production into a water phase and an oil phase can include adding 1 part-per-million (ppm) to 5 10000 ppm of the N-vinylpyrrolidone-based cationic copolymer to the oil-in-water emulsion, based on the total volume of the oil-in-water emulsion, to form a water phase and an oil phase. As used herein, ppm is a concentration where one ppm is equivalent to one part per 1000000 parts (e.g. 1 microliters N-vinylpyrrolidone-based cationic copolymer per liter of oil-in-water emulsion). Other examples of separating the oil-in-water emulsion formed during crude oil 10 production into a water phase and an oil phase can include adding 10 ppm to below 10000 ppm, adding 10 ppm to 1000 ppm, or adding 10 ppm to 100 ppm of the N-vinylpyrrolidone-based cationic copolymer to the oil-in-water emulsion, based on the total volume of the oil-in-water emulsion, to form a water phase and an oil phase.

Once formed, the water phase is separated from the oil phase. Once separated, either one 15 of the water phase and/or the oil phase can be recovered for further processing. It may be possible that the oil phase so produced may be a dehydrated oil as is known in the art. For the various embodiments, it is also possible that the water phase might have a maximum of 15 ppm to 50 ppm of the particles (e.g., oil droplets). To determine the ppm of oil in the water phase use a standard oil ppm curve. To prepare the standard oil ppm curve, prepare a series of known 20 concentrations of oil (the same oil present in the water phase) in a solvent (e.g., toluene; 1,1,1-trichloroethane; or Freon) and test the samples using a visible or IR spectrometer. Prepare the standard oil ppm curve from the results of the test. Use the same solvent to extract oil from the water phase discussed herein. After the extraction, test the solvent using the visible or IR spectrometer in the same manner and compare the results to standard oil ppm curve. The ppm of 25 the particles (e.g., oil droplets) in the water phase can then be interpolated from the standard oil ppm curve. The method of the present disclosure includes adding the N-vinylpyrrolidone-based cationic copolymer to the oil-in-water emulsion, based on the total volume of the oil-in-water emulsion, to form a water phase and an oil phase. In the oilfield, process conditions and the location of chemical injection points for water treatment chemicals vary from site to site as 30 described in Arnold, K.; Stewart, M. Surface Production Operations; 3rd ed.; Elsevier/Gulf: Boston, 2008, Chapter 9, pp482-609 as well as Manning, F. S.; Thompson, R. E. Oilfield

Processing Volume 2: Crude Oil; Pennwell: Tulsa, 1995, Chapter 8, pp 145-158. The water phase of the oil-in-water emulsion can have a pH value in a range of 5 to 9. The N-vinylpyrrolidone-based cationic copolymer of the present disclosure can help to destabilize the oil-in-water emulsion so as to enhance flocculation and eventual coalescence of the dispersed phase. A mixing process can be used with the oil-in-water emulsion in breaking the emulsion with the N-vinylpyrrolidone-based cationic copolymer of the present disclosure. For example, sufficient agitation can be used to allow the N-vinylpyrrolidone-based cationic copolymer of the present disclosure to mix thoroughly with the oil-in-water emulsion, followed by a period of flow inside a separator to promote gravity separation. The process also requires sufficient retention time in the separators to resolve the oil and water phases. The process may also require the addition of heat, gas flotation, and coalescers to facilitate separating the emulsion.

As appreciated, the efficacy of the cationic copolymer (e.g., the N-vinylpyrrolidone-based cationic copolymer) of the present invention can be dependent upon a number of factors such as the properties of the crude oil and/or the water of the emulsion, the mixer type, and the design and operating conditions of the separating equipment. The most effective conditions for the separation may be at least partially determined through the use of a bottle testing procedure, as is known.

Other factors that can influence the separation can include, but are not limited to, temperature, pH, type of crude oil, brine composition, solids content, oil content, system residence time, and droplet size distribution. An increase in temperature can result in a decrease in emulsion stability. The pH of the oil-in-water emulsion may also affect the performance of the cationic copolymer of the present disclosure. Surprisingly, the N-vinylpyrrolidone-based cationic copolymer of the present disclosure is a high pH tolerant cationic water clarifier, which may offer differentiated solutions to those seeking to resolve oil-in-water emulsions having a high pH (e.g., those having a pH of 7.0 to 9.0).

Additionally, other additives such as conventional coagulants, conventional flocculants, alum, preservatives or a combination thereof may also be utilized with the N-vinylpyrrolidone-based cationic copolymer.

The following examples are presented to describe preferred embodiments and utilities of the invention and are not meant to limit the invention unless otherwise stated in the claims

appended hereto.

Examples

Example 1: A copolymer of N-vinyl-2-pyrrolidone (67 wt.%) and 2-ethyl-5 dimethylaminoethyl methacrylate ethyl sulfate (33 wt.%) (LUVIQUAT® PQ 11, BASF, 20 wt.% aqueous solution, weight average molecular weight approximately 1000000 grams/mole (g/mol)).

Comparative Example A: A copolymer of acrylamide (80 wt.%) and diallyldimethyl ammonium chloride (20 wt.%) (polyAAM-DADMAC, Aldrich #409081, 10 wt.% aqueous 10 solution).

Comparative Example B: A copolymer of N-vinyl-2-pyrrolidone (70 wt.%) and 3-methyl-N-vinylimidazolium chloride (30 wt.%) (LUVIQUAT® FC370, BASF, 40 wt.% aqueous solution, weight average molecular weight approximately 100000 g/mol).

Comparative Example C: A copolymer of N-vinyl-2-pyrrolidone (50 wt.%) and 3-methyl-N-vinylimidazolium chloride (50 wt.%) (LUVIQUAT® FC 550, BASF, 40 wt.% 15 aqueous solution, weight average molecular weight approximately 80000 g/mol).

Comparative Example D: A copolymer of N-vinyl-2-pyrrolidone (55 wt.%) and 3-methyl-N-vinylimidazolium chloride (45 wt.%) (LUVIQUAT® HM 552, BASF, 20 wt.% aqueous solution, weight average molecular weight approximately 400000 g/mol).

20 Aqueous Solution of the Cationic Polymer: 0.1 wt.% of active solution was prepared by dissolving appropriate amounts of above cationic polymer (Example 1 and Comparative Examples A-D) in 100 milliliter (mL) of deionized (DI) water.

Bottle Test Procedure

25 Bottle Test Procedure 1 –Synthetic Produced Oil-in-Water Emulsion pH about 6.9

Prepare a synthetic produced oil-in-water emulsion by mixing at room temperature (23 °C) 6.5 mL of mid-gravity Middle Eastern crude oil with 650 mL of DI water (1 wt.%) for about 10 seconds under high shear (12000 rpm) in a high shear rotor-stator mixer and continue the agitation for a further 4 minutes under high shear of 12000 rpm. The resultant synthetic 30 produced oil-in-water emulsion has a pH of about 6.9.

Distribute the synthetic produced oil-in-water emulsion into 6 reagent bottles (100 mL each). An image of the untreated bottles is captured immediately (Fig. 1A).

Add a dose (as indicated in the brief description of Figs. 1B and 1C) of a 0.1 wt.% aqueous solution of each cationic polymer (Example 1 and Comparative Examples A-D) to a separate bottle and include one "BLANK" bottle without any cationic polymer additive for comparison. The dosage level, in ppm (by weight), is based on the amount of cationic polymer added to the total weight of the reverse emulsion in each bottle. Shake all bottles by hand 50 times and allow to resolve for one minute to allow floc formation before capturing an image. Obtain images in the presence of a 100 ppm dosage of cationic polymer (Fig. 1B) and a 300 ppm dosage of cationic polymer (Fig. 1C).

It was observed that the resultant oil flocs separate and float on the surface of the water. It is evident from the results that the cationic polymer from Example 1 facilitates the separation of residual oil and solid particles from the synthetic produced oil-in-water emulsion at a concentration of both 100 ppm and 300 ppm, at a pH of about 6.9.

Bottle Test Procedure 2 – Synthetic Produced Oil-in-Water Emulsion pH about 7.7

Prepare the synthetic produced oil-in-water emulsion as in Bottle Test Procedure 1 except add 250 μ L of 2 wt.% aqueous NaOH solution to 650 mL of DI water before adding the crude oil. The resultant synthetic produced oil-in-water emulsion has a pH of about 7.7.

Distribute the synthetic produced oil-in-water emulsion into 6 reagent bottles (100 mL each). An image of the untreated bottles is captured immediately (Fig. 2A).

Add a dose (as indicated in the brief description of Figs. 2B and 2C) of a 0.1 wt.% aqueous solution of each cationic polymer (Example 1 and Comparative Examples A-D) to a separate bottle and include one "BLANK" bottle without any cationic polymer additive for comparison. The dosage level, in ppm (by weight), is based on the amount of cationic polymer added to the total weight of the reverse emulsion in each bottle. Shake all bottles by hand 50 times and allow to resolve for one minute to allow floc formation before capturing an image. Obtain images in the presence of a 50 ppm dosage of cationic polymer (Fig. 2B) and a 100 ppm dosage of cationic polymer (Fig. 2C).

It was observed that the resultant oil flocs separate and float on the surface of the water. It is evident from the results that the cationic polymer from Example 1 facilitates the separation

of residual oil and solid particles from the synthetic produced oil-in-water emulsion at a pH of about 7.7 at concentrations of both 50 ppm and 100 ppm.

Bottle Test Procedure 3 – Synthetic Produced Oil-in-Water Emulsion pH about 8.9

5 Prepare the synthetic produced oil-in-water emulsion as in Bottle Test Procedure 1 except add 300 μ L of 2 wt.% aqueous NaOH solution to 650 mL of DI water before adding the crude oil. The resultant synthetic produced oil-in-water emulsion has a pH of about 8.9.

Distribute the synthetic produced oil-in-water emulsion into 6 reagent bottles (100 mL each). An image of the untreated bottles is captured immediately (Fig. 3A).

10 Add a dose (as indicated in the brief description of Figs. 3B and 3C) of a 0.1 wt.% aqueous solution of each cationic polymer (Example 1 and Comparative Examples A-D) to a separate bottle and include one “BLANK” bottle without any cationic polymer additive for comparison. The dosage level, in ppm (by weight), is based on the amount of cationic polymer added to the total weight of the reverse emulsion in each bottle. Shake all bottles by hand 50
15 times and allow to resolve for one minute to allow floc formation before capturing an image. Obtain images in the presence of a 50 ppm dosage of cationic polymer (Fig. 3B) and a 100 ppm dosage of cationic polymer (Fig. 3C).

It was observed that the resultant oil flocs separate and float on the surface of the water. It is evident from the results that the cationic polymer from Example 1 facilitates the separation
20 of residual oil and solid particles from the synthetic produced oil-in-water emulsion at a pH of about 8.9 at concentrations of both 50 ppm and 100 ppm.

We claim:

1. A method of separating an oil-in-water emulsion formed during crude oil production into a water phase and an oil phase, comprising:
 - adding 1 part-per-million (ppm) to 10000 ppm of an N-vinylpyrrolidone-based cationic copolymer to the oil-in-water emulsion, based on a total volume of the oil-in-water emulsion, to form the water phase and the oil phase; and
 - separating the water phase from the oil phase.
2. The method of claim 1, where the N-vinylpyrrolidone-based cationic copolymer is formed with monomers of N-vinyl-2-pyrrolidone and a quaternized dimethylaminoethyl methacrylate.
3. The method of claim 2, where the quaternized dimethylaminoethyl methacrylate is 2-ethyldimethylaminoethyl methacrylate ethyl sulfate.
4. The method of claim 2, where the N-vinylpyrrolidone-based cationic copolymer is poly[N-vinyl-2-pyrrolidone-co-quaternized dimethylaminoethyl methacrylate].
5. The method of claim 4, where the quaternized dimethylaminoethyl methacrylate is 2-ethyldimethylaminoethyl methacrylate ethyl sulfate.
6. The method of claim 2, where the N-vinylpyrrolidone-based cationic copolymer is formed as a reaction product of about 67 weight percent (wt.%) of N-vinyl-2-pyrrolidone and about 33 wt.% of a quaternized dimethylaminoethyl methacrylate.
7. The method of claim 6, where the quaternized dimethylaminoethyl methacrylate is 2-ethyldimethylaminoethyl methacrylate ethyl sulfate.
8. The method of claim 1, where the N-vinylpyrrolidone-based cationic copolymer has a weight average molecular weight of about 1000000 g/mol.

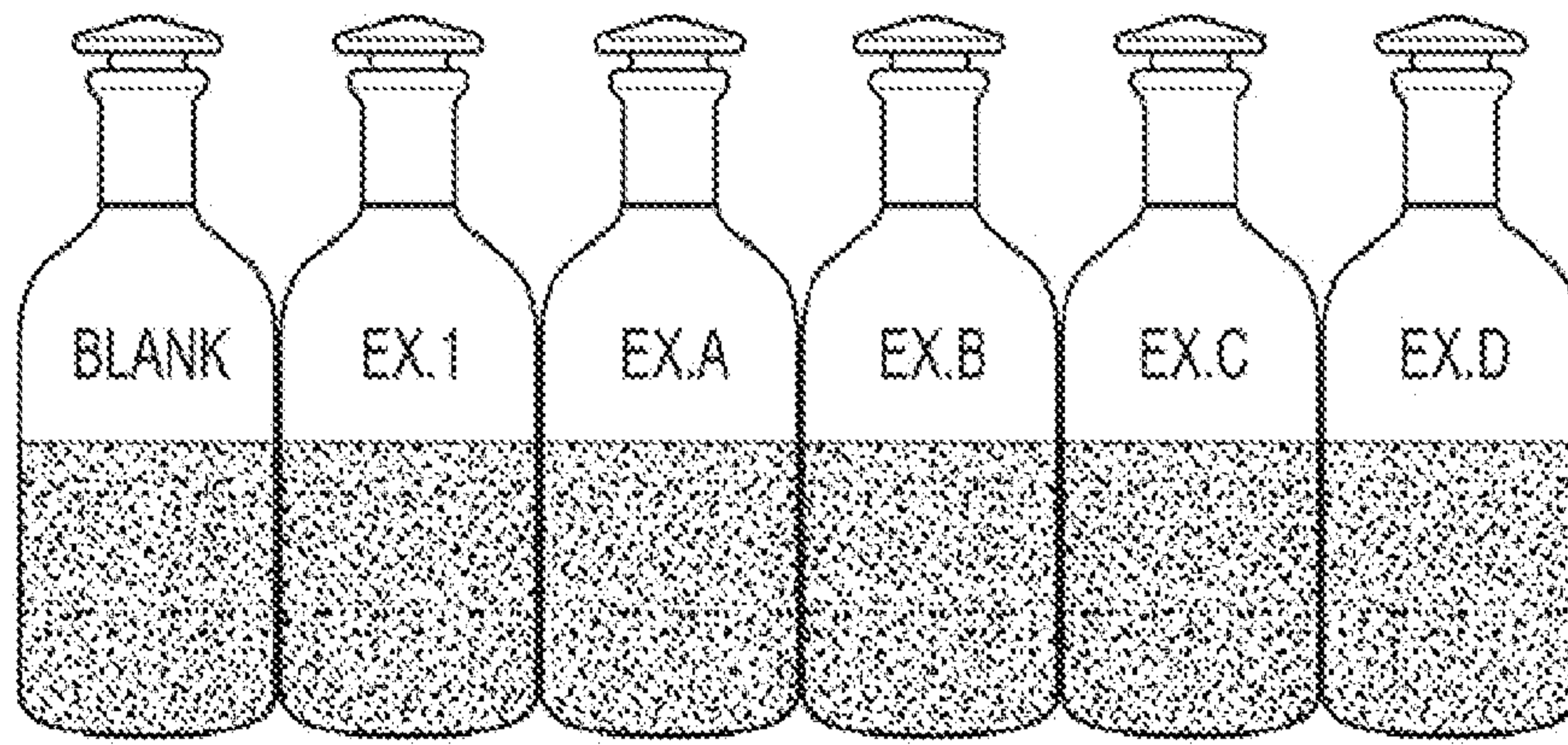


FIG. 1A

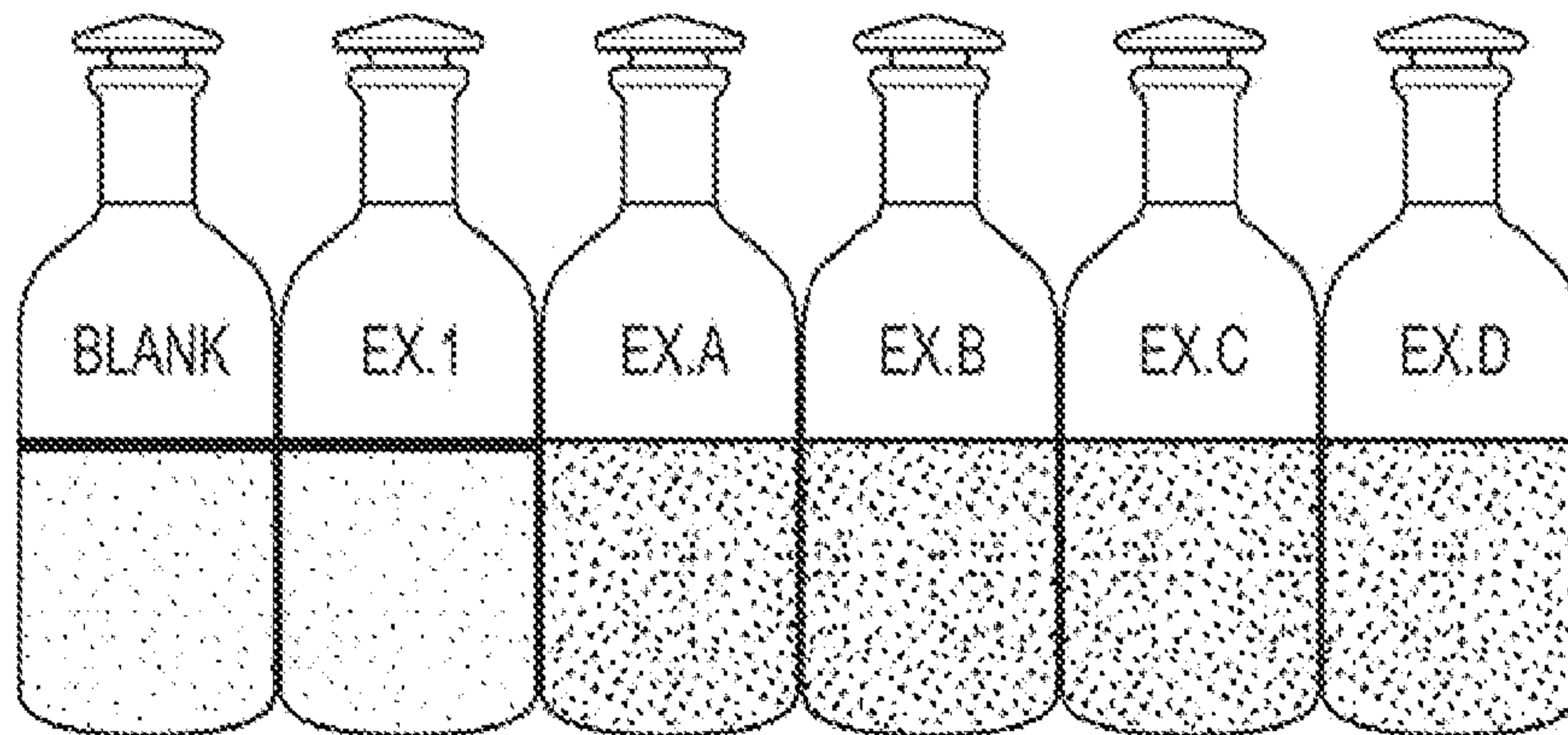


FIG. 1B

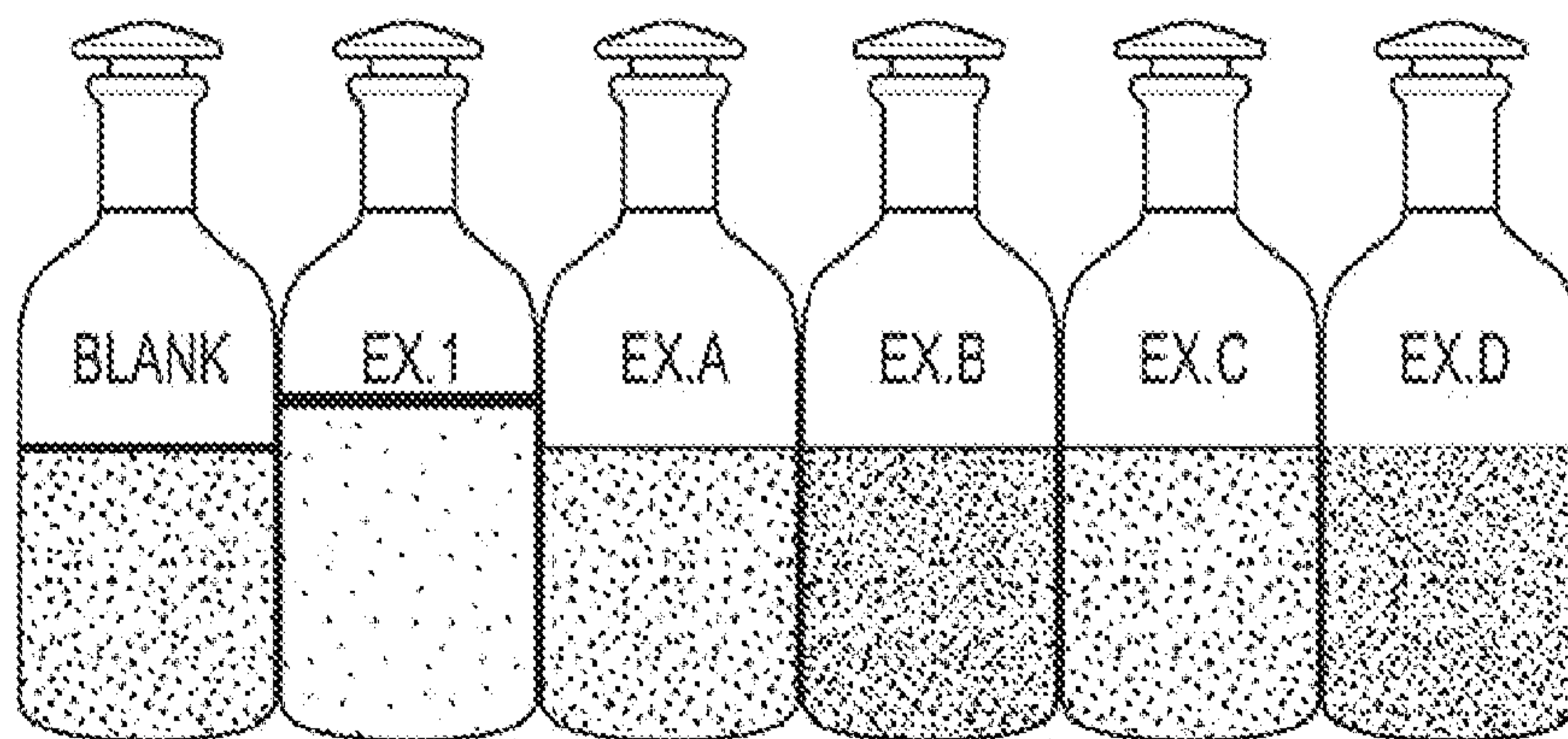


FIG. 1C

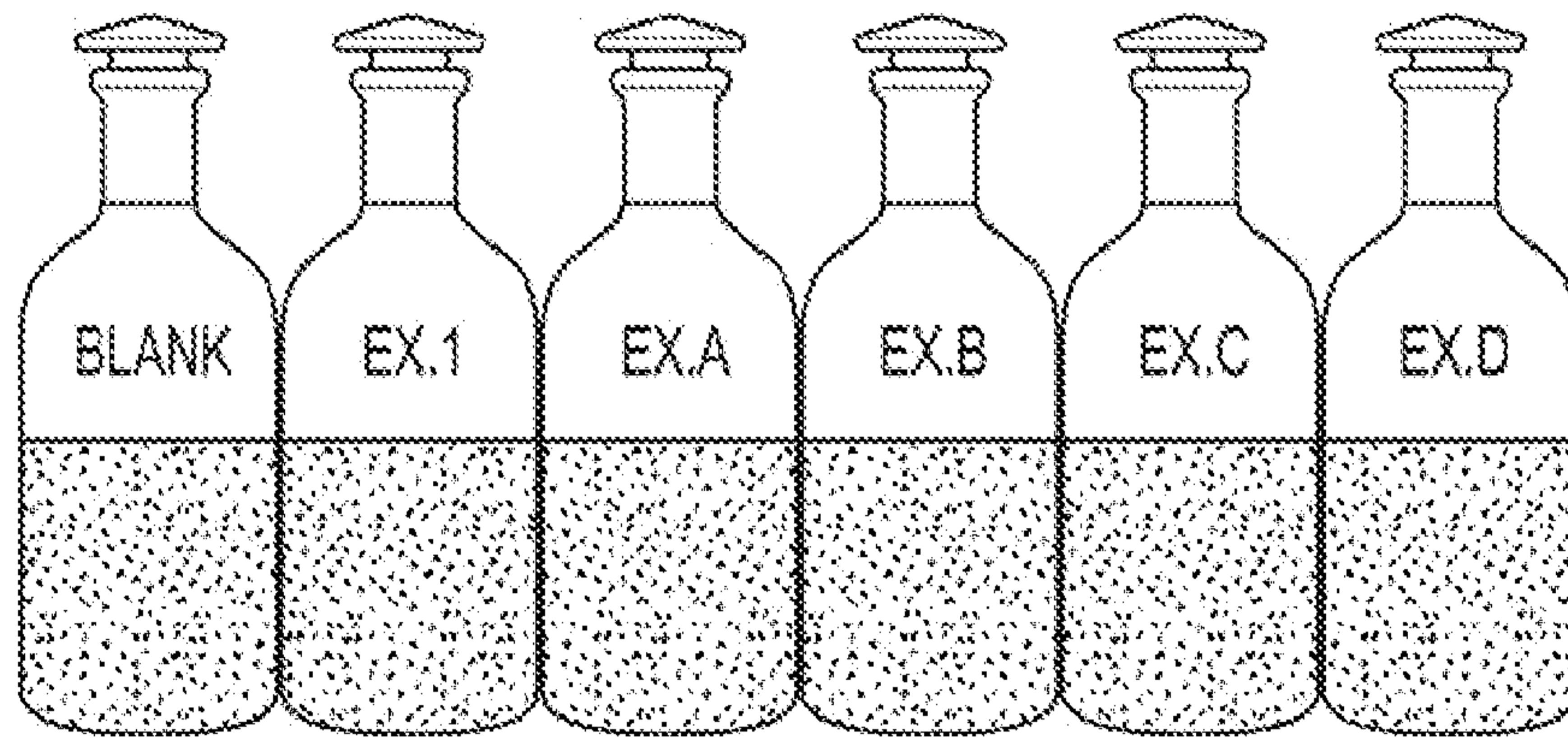


FIG. 2A

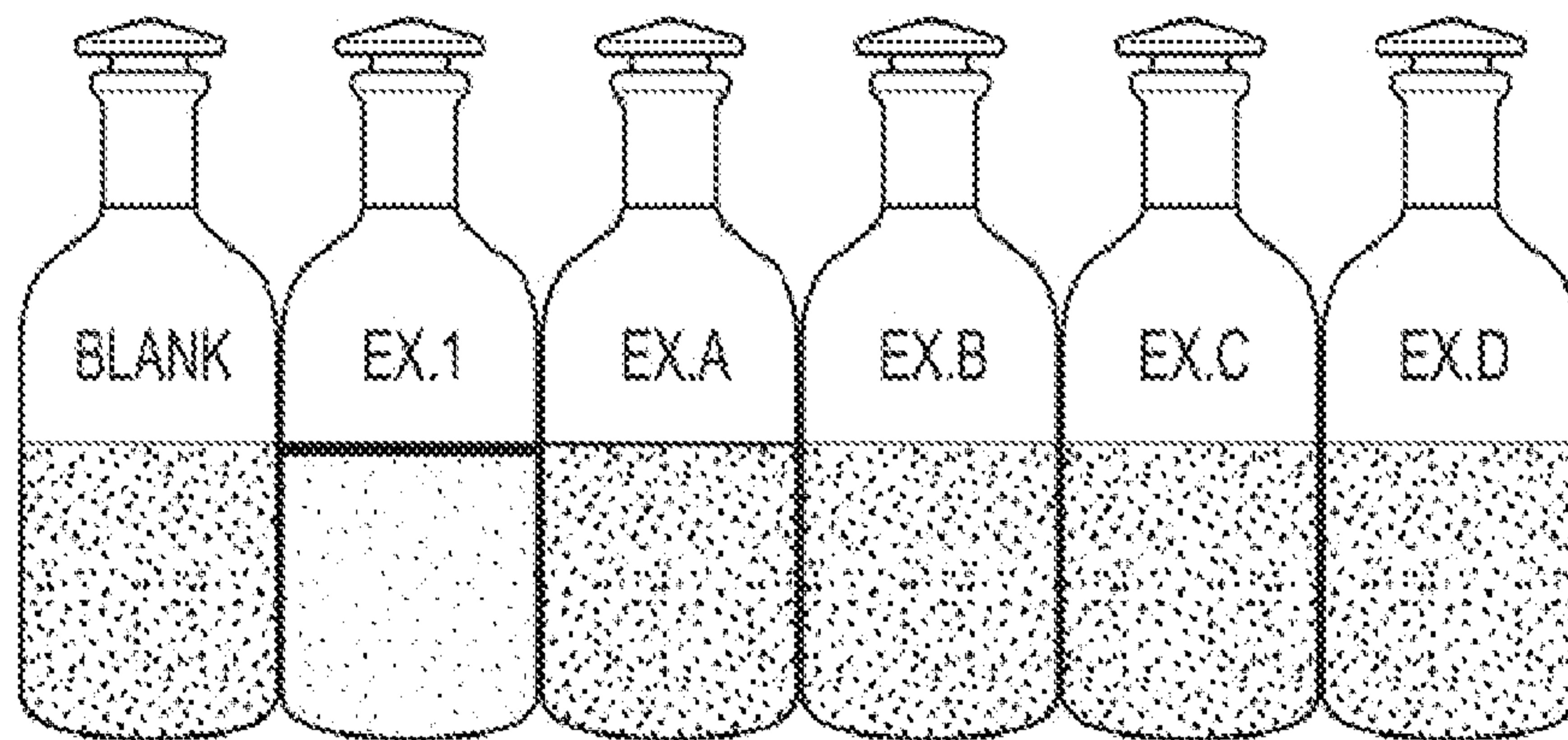


FIG. 2B

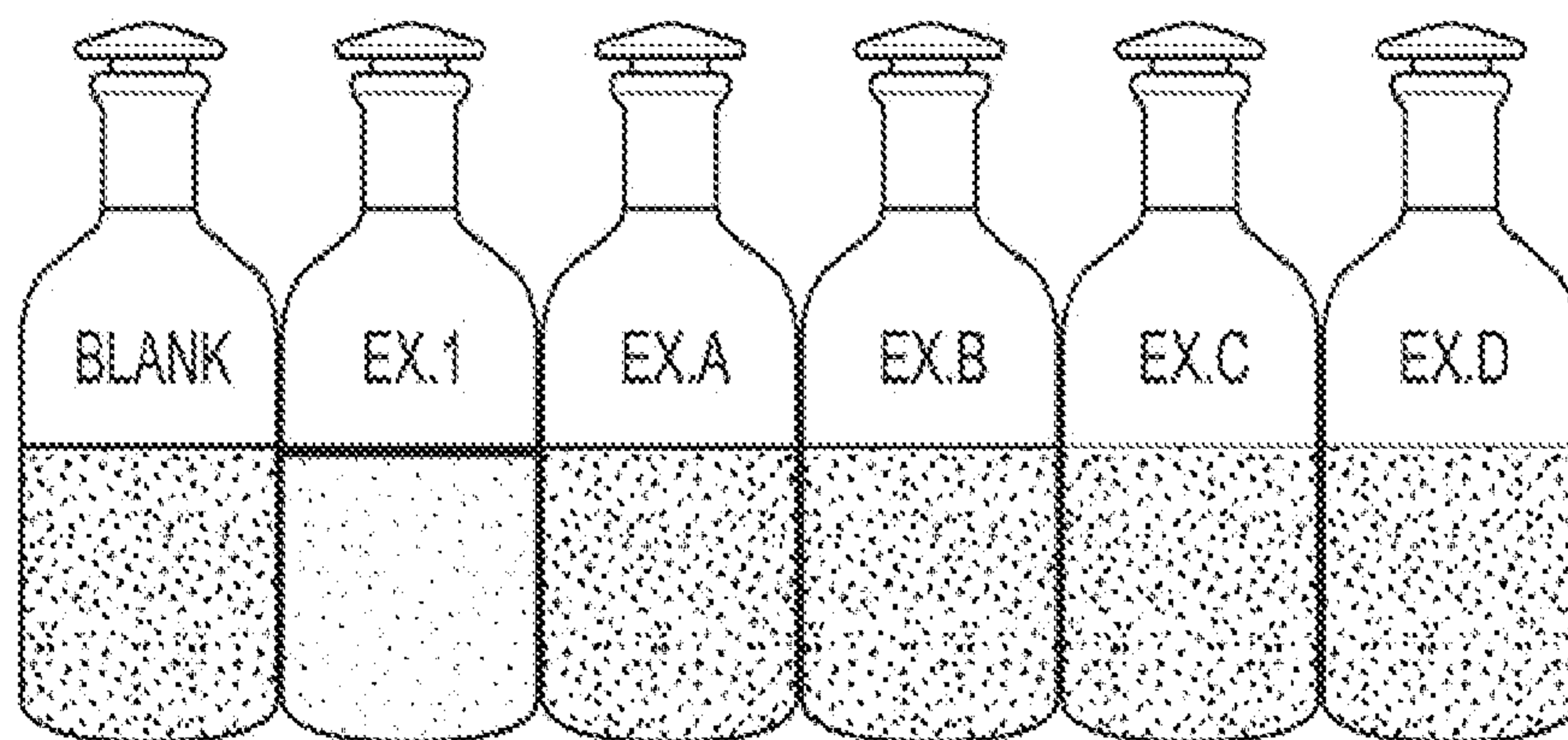


FIG. 2C

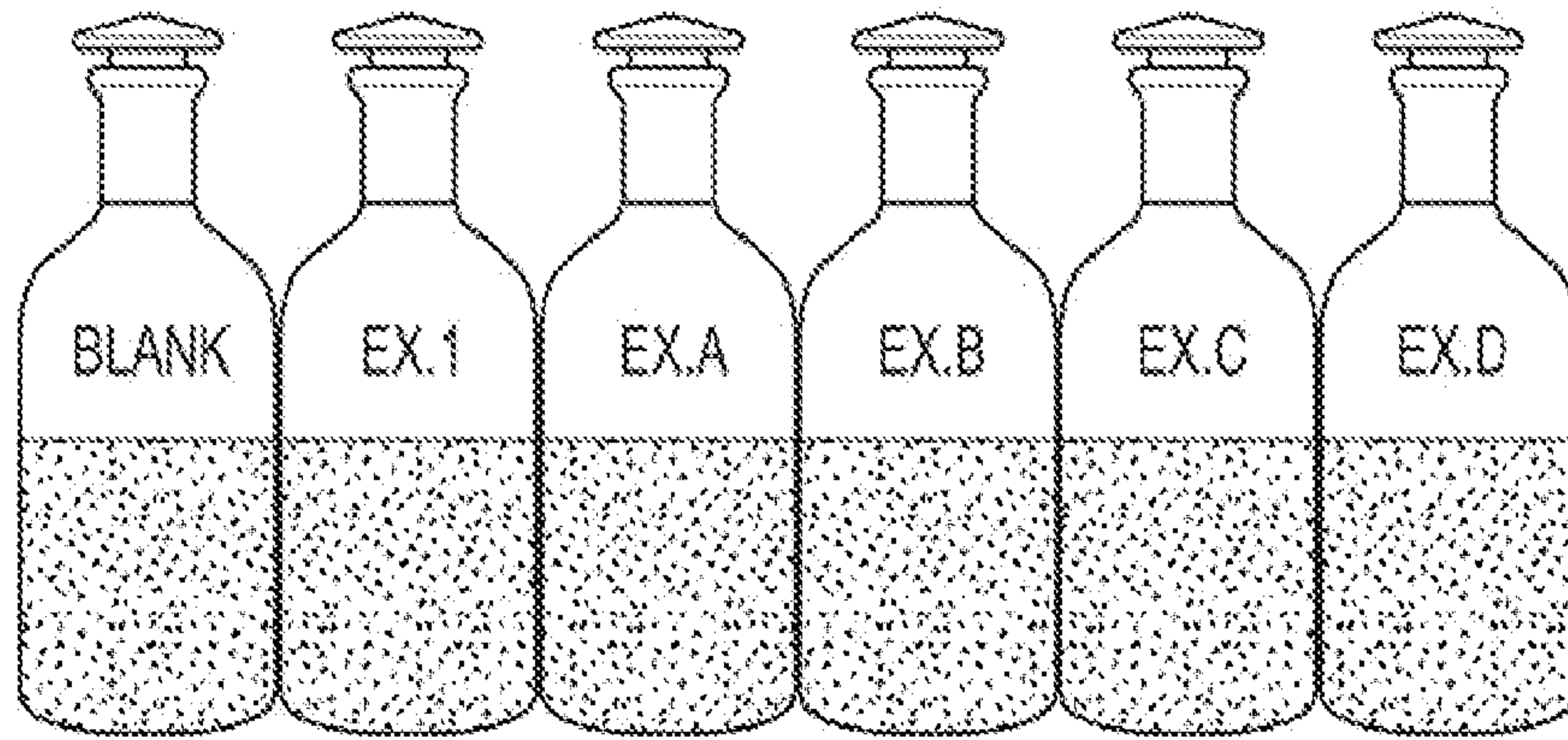


FIG. 3A

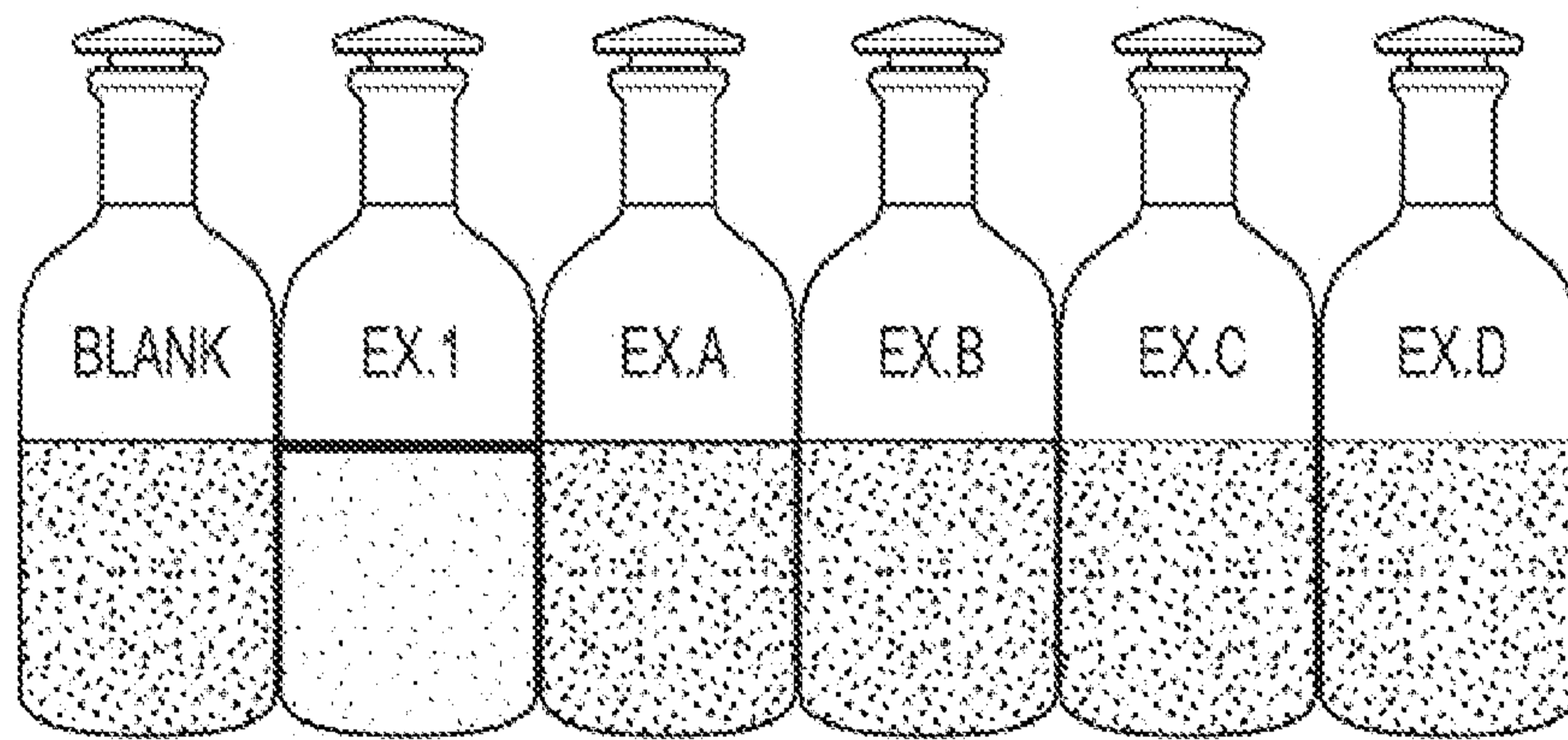


FIG. 3B

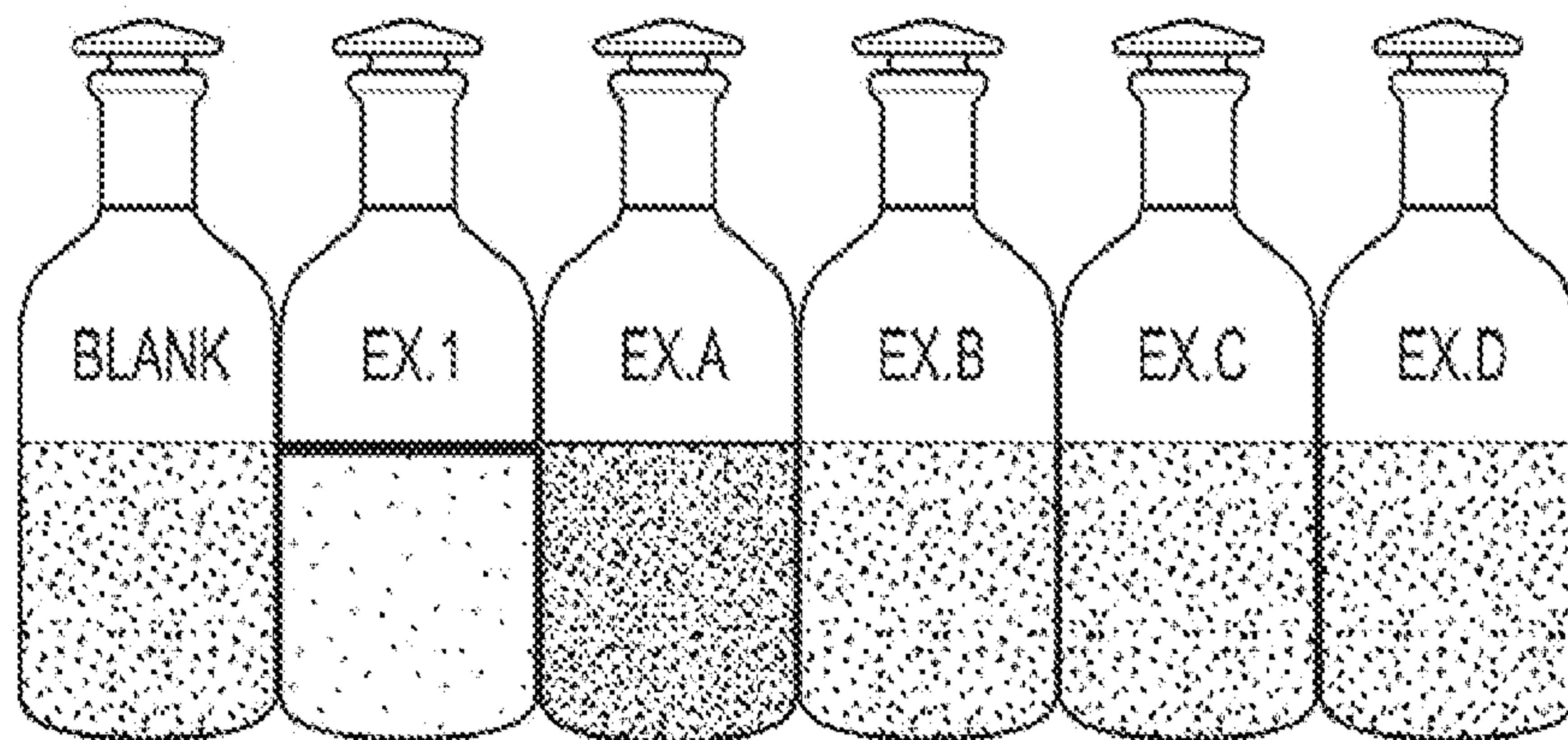


FIG. 3C

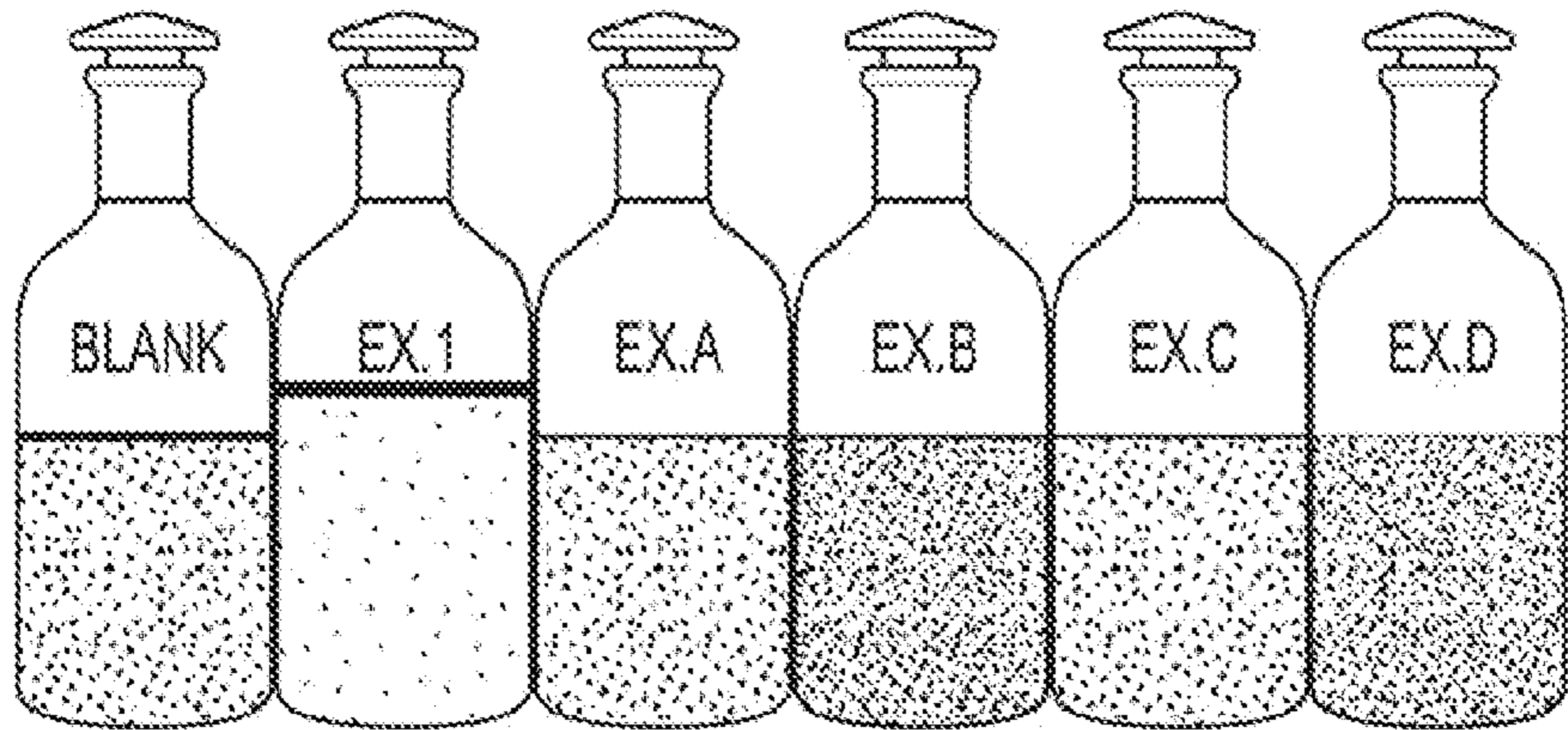


FIG. 1C