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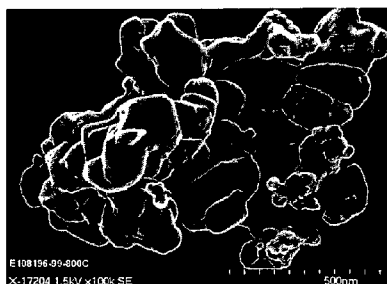
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(54) Title: PROCESSES FOR THE FLUX CALCINATION PRODUCTION OF TITANIUM DIOXIDE

(57) Abstract: Processes for the production of rutile titanium dioxide from titanyl hydroxide using calcination with a flux are provided. Calcination in the presence of sodium chloride flux has been found to lower the calcination temperature used to produce the rutile form of titanium dioxide and to improve control of particle size.

(a)



(b)





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- *before the expiration of the time limit for amending the claims and to be republished in the event of receipt of amendments*

TITLE

PROCESSES FOR THE FLUX CALCINATION PRODUCTION OF TITANIUM DIOXIDE

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FIELD OF THE INVENTION

The present invention relates to processes for the production of rutile titanium dioxide from titanyl hydroxide using calcination with a flux. Titanium dioxide, particularly the rutile phase, is used as a white pigment in paints and plastics.

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BACKGROUND

Titanyl hydroxide can be produced by two major processes, chloride and sulfate. Calcination in the presence of sodium chloride flux lowers the calcination temperature used to produce the rutile form of titanium dioxide. Robert (US 5494652) discloses a process reacting tin oxide with an alkali metal halide at 400 to 1200 °C.

Disclosed herein is a calcination process using a sodium chloride flux for the production of titanium dioxide. The specific parameters of the process produce the rutile phase of titanium dioxide. The process can produce pigmentary-sized rutile.

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SUMMARY OF THE INVENTION

One aspect of the present invention is a process for producing titanium dioxide comprising:

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a) mixing sodium chloride with titanyl hydroxide and, optionally, one or more morphology control additives, to form a mixture;

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b) heating the mixture to a target temperature of 800 to 1200 °C to form a mixture of titanium dioxide and sodium chloride; and

c) optionally, separating the sodium chloride from the titanium dioxide.

In some embodiments, the heating is carried out over a time period of about 0.5 hours to about 48 hours.

In some embodiments, the mixture is held at the target temperature for up to 72 hours.

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BRIEF DESCRIPTION OF THE FIGURES

Figure 1 (a) is a scanning electron micrograph of irregularly-shaped particles with a size range of about 50 to 300 nm.

Figure 1 (b) is a scanning electron micrograph of well-shaped particles with a size range of about 200 to 800 nm, and illustrates how NaCl can serve as a size and shape control agent.

Figure 2 (a) is a scanning electron micrograph showing media-milled product mainly of 20-100 nm irregularly-shaped particles.

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Figure 2 (b) is a scanning electron micrograph showing media-milled product of well-shaped primary particles in the range of about 100-500 nm.

DETAILED DESCRIPTION

Flux calcination crystallization using sodium chloride involves conversion of amorphous titanyl hydroxide to the rutile form of titanium dioxide at relatively low temperature conditions (as low as 800 °C) compared to the calcination temperatures without the addition of sodium chloride (ca. 1000 °C) typically utilized in commercial titanium dioxide production. The titanyl hydroxide starting material can be produced by the commercially known sulfate or chloride processes or by other processes. Reaction temperatures in the flux calcination crystallization process range from as low as 800 °C up to 1200 °C. Reaction

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times range from a fraction of a minute to three days. The specific structure-directing flux, sodium chloride, can be used to control the production of the rutile structural form of titanium dioxide. Variation of the range of process conditions such as control of the time at temperature in the reaction mixture can be used to
5 selectively control the resulting titanium dioxide particle size and morphology.

The rutile phase of titanium dioxide of pigmentary size can be formed at 800 °C.

10 In a process of the present invention for the production of rutile titanium dioxide, titanyl hydroxide is mixed with sodium chloride. Titanyl hydroxide can be produced by either of the known commercial processes for titanium dioxide production, the chloride process or the sulfate process. Additionally, titanyl hydroxide can be produced by other known processes such as extraction of
15 titanium-rich solutions from digestion of ilmenite by oxalic acid or hydrogen ammonium oxalate.

After mixing the titanyl hydroxide with sodium chloride, the resulting mixture is heated to a target temperature of 800 to 1200 °C to form titanium
20 dioxide. In preferred embodiments, the heating is carried out over a time period of about 0.5 hours to about 48 hours. In some embodiments, the mixture is held at the target temperature for up to 72 hours. The process produces a product comprising titanium dioxide and some of the starting sodium chloride. If desired, the amount of sodium chloride in the product can be reduced by washing or by
25 other separation techniques such as vacuum distillation at about 1000 °C.

The concentration of the sodium chloride in the mixture before heating is a factor in controlling the resulting primary particle size and degree of agglomeration and aggregation, i.e., the secondary particle size, of the titanium
30 dioxide obtained from the process. The processes disclosed herein can produce pigmentary-sized titanium dioxide. An average particle diameter of 100 nanometers is usually used to divide nano-sized titanium dioxide from pigmentary-sized titanium dioxide. 100 nanometers is at the low end of the size

range of pigmentary titanium dioxide supplied by the existing commercial processes. Smaller particle diameters are referred to as nano-sized titanium dioxide. Pigmentary-sized particles have a large market and thus are frequently the desired particle size. The time at temperature is an important factor in
5 determining the particle size of the resulting titanium dioxide with increasing time at temperature leading to increasing particle size. Titanium dioxide is frequently supplied to the pigment market with a coating such as aluminum which can be added in an additional process step.

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EXAMPLES

Example 1

This example illustrates the use of NaCl to control the morphology of
rutile.

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96.0 g ammonium titanyl oxalate (ATO), Aldrich 99.998, were dissolved in 400 mL deionized water and the resulting mixture was filtered to remove undissolved solids. The filtered solution was transferred to a jacketed Pyrex round-bottomed flask equipped with a water-cooled condensor and heated to
20 90°C with stirring using a Teflon-coated stirring bar. A solution consisting of 1 part concentrated NH_4OH and 1 part deionized water by volume was added dropwise to the ATO solution until a pH of 7.5 was attained. The white slurry was stirred at 90°C for 15 minutes after which time it was transferred to a jacketed filter filtered at 90°C. The filter cake was washed several times with water heated
25 to 90°C until the filtrate had a conductivity of about 500 microSiemens. A small portion of the washed cake was dried in air at room temperature. X-ray powder diffraction showed the dried sample to be nanocrystalline anatase.

Another portion of dried sample was heated in air from room temperature
30 to 800°C over a time period of 3 hours, and held at 800°C for 1 hour. XPD showed the fired product to consist of 99% rutile and 1% anatase. A scanning electron micrograph of this sample (Figure 1(a)) shows irregularly-shaped particles with a size range of about 50 to 300 nm.

Another portion of dried sample was mixed with NaCl by grinding in a mortar. The amount of NaCl was 5 wt% based on the weight of dry TiO₂. The mixture was heated in air from room temperature to 800°C over a time period of 3 hours, and held at 800°C for 1 hour. XPD showed the fired product to consist
5 mainly of rutile with a trace of Na₂Ti₆O₁₃. No anatase was found. A scanning electron micrograph of this sample (Figure 1 (b)) shows well-shaped particles with a size range of about 200 to 800 nm, and illustrates how NaCl can serve as a size and shape control agent.

10 Another portion dried sample was mixed with NaCl by grinding in a mortar. The amount of NaCl was 5 wt% based on the weight of dry TiO₂. The mixture was heated in air from room temperature to 850°C over a time period of 3 hours, and held at 850°C for 1 hour. XPD showed the fired product to consist
15 mainly of rutile with a trace of Na₂Ti₆O₁₃. No anatase was found.

Example 2

This example illustrates the use of NaCl as a rutile promoter.

20 2.5 g samples of titanyl hydroxide, derived from an oxalate process leachate and containing about 0.5 g TiO₂ on a dry basis, were heated with NaCl (5 wt% and 33 wt% on TiO₂), and with AlCl₃·6H₂O (1 wt% Al₂O₃ on TiO₂), as described in Table 1 below, in an alumina crucible from room temperature to 90°C over a 1 hour period and held at 90°C for 4 hours at which time the
25 temperature was increased to 850°C over a 3 hour period, and held at 850°C for 1 hour. Results of X-ray powder diffraction analyses are given in the Table and indicate that NaCl greatly assists the formation of rutile, while in the absence of NaCl, anatase is the predominant product. The results also show addition of aluminum chloride counteracts the sodium chloride and stabilizes anatase.

Table 1

Ingredient (g)	A	B	C	D	E
Ti-ppt	2.5	2.5	2.5	2.5	2.5
NaCl	-	0.025	0.167	0.025	0.167
AlCl ₃ ·6H ₂ O	-	-	-	0.024	0.024
H ₂ O	1	1	1	1	1
Product	anatase	anatase + tr. Na ₂ Ti ₆ O ₁₃	~98% rutile ~2% anatase tr. Na ₂ Ti ₆ O ₁₃	anatase	anatase
Ingredient (g)	A	B	C	D	E
Ti-ppt	2.5	2.5	2.5	2.5	2.5
NaCl	-	0.025	0.167	0.025	0.167
AlCl ₃ ·6H ₂ O	-	-	-	0.024	0.024
H ₂ O	1	1	1	1	1
Product	anatase	anatase + tr. Na ₂ Ti ₆ O ₁₃	~98% rutile ~2% anatase tr. Na ₂ Ti ₆ O ₁₃	anatase	anatase

The reaction of Example 2C was repeated without the initial four hour heating at 90°C and the reaction mixture was heated from room temperature to 850°C over a 3 hour period and held at 850°C for 1 hour. From XPD, the product was identified as mainly rutile with traces of anatase and Na₂Ti₆O₁₃.

Example 3

This example illustrates the use of NaCl as a rutile promoter.

A portion of titanyl hydroxide, derived from an oxalate process leachate, was dried in air at room temperature and used for experiments 3A and 3B (Table 2). 0.6 g samples, containing about 0.5 g TiO₂ on a dry basis, were ground together with NaCl (5 wt% and 33 wt% NaCl based on TiO₂) and heated in alumina crucibles from room temperature to 850°C over a 3 hour period, and held at 850°C for 1 hour. Results of X-ray powder diffraction analyses are given in Table 2 and indicate that NaCl greatly assists the formation of rutile.

Table 2

Ingredient (g)	3A	3B
Ti-ppt (g)	0.6	0.6
NaCl (g)	0.025	0.167
Product	~ 1:1 rutile:anatase Trace $\text{Na}_2\text{Ti}_6\text{O}_{13}$	Essentially all rutile with tr. $\text{Na}_2\text{Ti}_6\text{O}_{13}$

Example 4

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This example shows that NaCl is a rutile promoter when particle size control additives used in the sulfate process are also present.

A portion of titanyl hydroxide, derived from an oxalate process leachate,
 10 was dried in air at room temperature and used for experiments Example 4A and
 Example 4B. 0.6 g samples, containing about 0.5 g TiO_2 on a dry basis, were
 ground together with 0.0005 g Na_2SO_4 , 0.0025 g K_2SO_4 , 0.0024 g, $\text{NH}_4\text{H}_2\text{PO}_4$,
 and 0.025 g rutile seed. 0.025 g NaCl (5 wt%) were added to sample B and both
 15 samples were heated in alumina crucibles from room temperature to 800°C over
 a 3 hour period, and held at 800°C for 1 hour. Results of X-ray powder
 diffraction analyses are given in Table 3 and indicate that NaCl greatly assists the
 formation of rutile.

Table 3

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	Example 4 A	Example 4 B
Product:	Mainly anatase with v. small amount rutile	~ 60% rutile, 40% anatase

Example 5

This example illustrates the use of NaCl to control the morphology of rutile.

5

36.8 g ammonium titanyl oxalate (ATO), Aldrich 99.998, were dissolved in 300 mL deionized water and the resulting mixture was filtered to remove undissolved solids. The filtered solution was transferred to a Pyrex beaker and stirred with a Teflon-coated stirring bar. Concentrated NH_4OH was added dropwise to the ATO solution until a pH of 9 was attained. The white slurry was filtered immediately and the filter cake was washed with 400 mL deionized water at room temperature. The Ti-containing cake was transferred to a beaker and 450 mL concentrated NH_4OH were added and the mixture was stirred and boiled for 30 minutes. The precipitate filtered rapidly. The Ti cake was again transferred to a beaker and reslurried with concentrated NH_4OH , then boiled for 30 minutes. After collecting the solids on a filter, the cake was transferred to a beaker, slurried with about 450 mL deionized water, stirred for one day at room temperature, then boiled for one hour. After collecting the solids, the washed cake was dried in air under IR heat ($\sim 40^\circ\text{C}$). The entire sample was heated to 800°C over a period of three hours, and held at 800°C for three hours. An X-ray powder diffraction pattern of the fired product showed it to be mainly rutile with a trace of anatase. Scanning electron microscopy imaging showed media-milled product to consist mainly of 20-100 nm irregularly-shaped particles as shown in Figure 2(a).

25

Ti-precipitate cake was made as described above, but before drying the washed cake under IR heat, 3.32 g NaCl, dissolved in 10 mL H_2O , were mixed into the TiO_2 cake. The entire sample was heated to 800°C over a period of three hours, and held at 800°C for one hour. An X-ray powder diffraction pattern of the fired product showed it to be 95% rutile and 5% anatase. Scanning electron microscopy imaging showed media-milled product to consist of well-shaped primary particles in the range of about 100-500 nm and some small, <100 nm, irregularly-shaped particles as shown in Figure 2(b).

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Example 6

This example shows that NaCl is a rutile promoter when particle size control additives used in the sulfate process are also present, and when the mixture is heated in a rotary calciner.

315 g of titanyl hydroxide, derived from an oxalate process leachate, containing ~ 49 g TiO_2 on a dry basis, was mixed with 49 g of a solution consisting of 0.19 wt% KH_2PO_4 , 0.38 wt% K_2HPO_4 , and 0.09 wt% Na_2HPO_4 in H_2O , 98 g of a solution consisting of 2.4 wt% $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ in H_2O , 49 g of a solution consisting of 4.8 wt% NaCl in H_2O , and 69 g of 2.9 wt% rutile seed suspension in aqueous HCl solution. The mixture was dried in air under IR heat (~ 40°C) and powdered in a mortar. 55 g of the dried mixture were heated to 1050°C in a fused silica rotary calciner over a period of 3 hours and held at 1050°C for 8 hours. An XPD pattern of the product showed it to be all rutile.

2.5 g of the dried mixture prepared as described above were fired in an alumina crucible to 1050°C over a 12 hour period at which point power to the furnace was removed and the sample was allowed to cool naturally to room temperature. An XPD pattern of the product showed it to be rutile with a trace of $\text{Na}_2\text{Ti}_6\text{O}_{13}$.

2.5 g of the dried mixture prepared as described above were fired in an alumina crucible to 1150°C over a 12 hour period at which point power to the furnace was removed and the sample was allowed to cool naturally to room temperature. An XPD pattern of the product showed it to be all rutile.

Example 7

Titanyl hydroxide, derived from an oxalate process leachate, was washed
 with water at room temperature to remove NH_4OH via cycles of stirring and
 centrifuging until the pH was about 7-8. The slurry used for the experiments
 contained 13.18 wt% TiO_2 as shown in Table 4. Phosphate, potassium and
 sodium additives were mixed with the titanyl hydroxide as indicated in Table 4.
 Sodium chloride flux was added to some of the mixtures. When sodium chloride
 was present, a greater amount of rutile was observed at the lower target
 temperatures, showing that NaCl is a good rutile promoter. SEM images showed
 that NaCl was a particle morphology control agent at 800°C.

Table 4

g of Washed Oxalate Process Leachate Ti- ppt., 13.2 wt% TiO_2	g of 0.48 wt% H_3PO_4 Solution	g of 0.43 wt% KCl Solution	g of 0.07 wt% NaCl Solution	g of 3.0 wt% NaCl Solution	Ramp-up time (hrs) to Target Temp	Target Temp °C	Hold Time (hrs) at Target Temp	% Rutile	% Anatase
15.2	1	3	1	2	5	1050	4	100	0
15.2	3	1	3	0	5	1050	4	99	0
15.2	1	3	3	2	5	850	0	60	40
15.2	3	1	1	0	5	850	0	0	100
15.2	3	1	3	2	5	1050	0	100	0
15.2	1	3	1	0	5	1050	0	96	4
15.2	1	1	1	2	15	1050	0	100	0
15.2	3	3	3	0	15	1050	0	100	0
15.2	3	1	1	2	5	850	4	99	1
15.2	1	3	3	0	5	850	4	4	96
15.2	1	1	3	0	15	850	0	0.5	99.5
15.2	3	3	1	2	15	850	0	100	0
15.2	3	3	3	2	15	1050	4	100	0
15.2	1	1	1	0	15	1050	4	100	0
15.2	1	1	3	2	15	850	4	100	0
15.2	3	3	1	0	15	850	4	3	97
15.2	1	0	3	0	15	800	0	0	100
15.2	1	1	3	3.33	15	800	0	12	88
15.2	1	0	3	3.33	15	850	0	100	0
15.2	1	0	3	0	15	800	0	0	100
15.2	1	1	3	3.33	16	800	4	100	0

CLAIMS

What is claimed:

1. A process for producing titanium dioxide comprising:
5
a) mixing sodium chloride with titanyl hydroxide and, optionally, one or more morphology control additives, to form a mixture;
b) heating the mixture to a target temperature of 800 to 1200
10 °C to form a mixture of titanium dioxide and sodium chloride; and
c) optionally, separating the sodium chloride from the titanium dioxide.
15
2. The process of claim 1 wherein the heating is carried out over a time period of about 0.5 hours to about 48 hours.
3. The process of claim 1 wherein the mixture is held at the target temperature for up to 72 hours.
20
4. The process of claim 1 wherein the amount of sodium chloride is from 1 wt% to 50 wt% based on the amount of titanium dioxide that can be recovered from the titanyl hydroxide.
- 25 5. The process of claim 1 wherein the amount of sodium chloride is from 1 wt% to 30 wt% based on the amount of titanium dioxide that can be recovered from the titanyl hydroxide.

6. The process of claim 1 wherein the amount of sodium chloride is from 1 wt% to 10 wt% based on the amount of titanium dioxide that can be recovered from the titanyl hydroxide.

5 7. The process of claim 1 wherein the sodium chloride is in solid form.

8. The process of claim 1 wherein the sodium chloride is in solution.

9. The process of claim 1 wherein the mixing comprises stirring, shaking,
10 or tumbling for several minutes up to several days.

10. The process of claim 1 wherein the morphology control additives contain ions selected from the group consisting of phosphate, potassium, sodium, and aluminum.

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11. The process of claim 1 wherein the titanyl hydroxide is produced from titanyl sulfate solution.

12. The process of claim 1 wherein the titanyl hydroxide is produced from
20 titanium oxychloride solution.

13. The process of claim 1 wherein the titanyl hydroxide is produced from ammonium titanyl oxalate solution.

25 14. The process of claim 1 wherein the heating is performed in a rotary calciner.

15. The process of claim 1 wherein the heating is performed in a static calciner.

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16. The process of claim 1 wherein the sodium chloride is separated from the titanium dioxide by water washing.

17. The process of claim 1 wherein the sodium chloride is separated from
5 the titanium dioxide by vacuum distillation.

18. The process of claim 1 wherein the titanium dioxide comprises the rutile form.

10 19. The process of claim 18 wherein the rutile titanium dioxide has a particle size of 50 to 1000 nm.

20. The process of claim 18 wherein the rutile titanium dioxide has a particle size of 50 to 500 nm.

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21. The process of claim 1 wherein the mixture is heated to a temperature of 800-1000 °C.

22. The process of claim 1 wherein the mixture is heated to a
20 temperature of 800-900 °C.

23. The process of claim 1 wherein the target temperature is reached within a time period of 0.5 to 48 hours.

25 24. The process of claim 1 wherein the target temperature is reached within a time period of 1 to 24 hours.

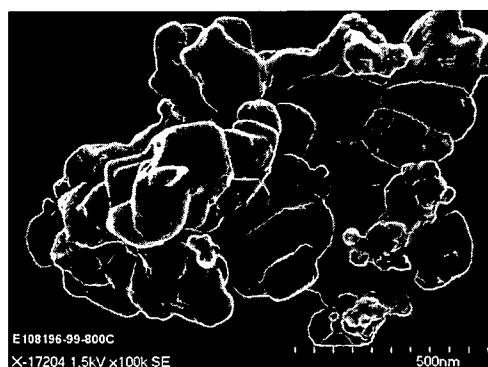
25. The process of claim 1 wherein the mixture is held at the target temperature for up to 72 hours.

30

26. The process of claim 1 wherein the mixture is held at the target temperature for up to 48 hours.

27. The process of claim 1 wherein the mixture is held at the target
5 temperature for up to 24 hours.

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Figure 1(a)**Figure 1(b)**

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Figure 2(a)

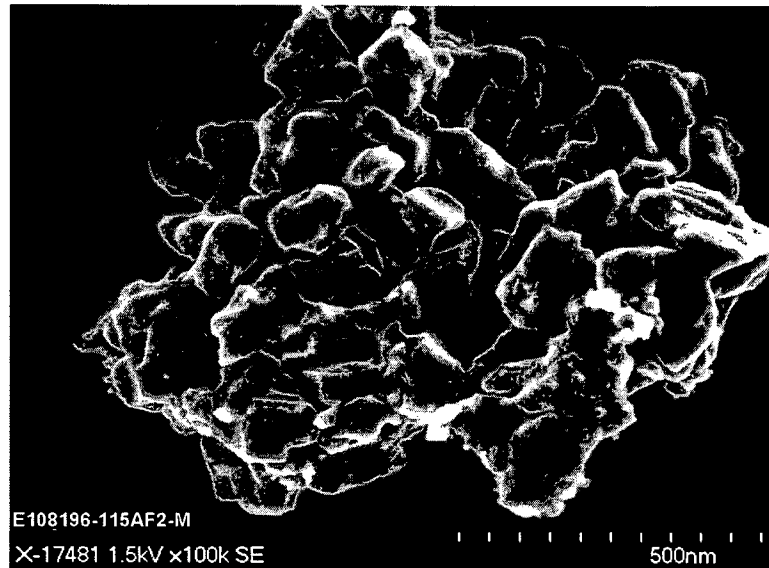


Figure (2b)



INTERNATIONAL SEARCH REPORT

International application No

PCT/US2007/026328

A. CLASSIFICATION OF SUBJECT MATTER
INV. C01G23/047 C01G23/08

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
C01G

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal, WPI Data

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	GB 1 204 601 A (DU PONT [US]) 9 September 1970 (1970-09-09) page 1, line 57 - page 2, line 16 page 3, line 11 - line 51 page 3, line 95 - line 100 page 4, line 20 - line 35 page 4, line 59 - line 108 page 5, line 42 - line 55; examples 3,5; table 1	1-27
X	US 3 728 443 A (BERISFORD R ET AL) 17 April 1973 (1973-04-17) column 2, line 17 - column 3, line 39 ----- -/--	1-3,7, 9-11,15, 16,18-27

☒ Further documents are listed in the continuation of Box C.

☒ See patent family annex.

* Special categories of cited documents:

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Date of the actual completion of the international search

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Name and mailing address of the ISA/

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INTERNATIONAL SEARCH REPORT

International application No

PCT/US2007/026328

C(Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 0 341 703 A (ISHIHARA MINING & CHEMICAL CO [JP]) 15 November 1989 (1989-11-15) column 3, line 20 - column 4, line 28; example 1 column 5, line 7 - line 25 -----	1-3,7,9, 10,15, 16,21, 22,25-27
X	US 5 494 652 A (ROBERT JEAN C [FR]) 27 February 1996 (1996-02-27) cited in the application column 2, line 5 - column 3, line 7 -----	1,4,16, 21,22

INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No

PCT/US2007/026328

Patent document cited in search report		Publication date		Patent family member(s)	Publication date
GB 1204601	A	09-09-1970	DE	1767021 A1	19-08-1971
			FI	47458 B	31-08-1973
			FR	1561551 A	28-03-1969
<hr/>					
US 3728443	A	17-04-1973	NONE		
<hr/>					
EP 0341703	A	15-11-1989	CA	1340573 C	01-06-1999
			DE	68908464 D1	23-09-1993
			DE	68908464 T2	05-01-1994
			JP	1286924 A	17-11-1989
			JP	1899158 C	23-01-1995
			JP	6024977 B	06-04-1994
<hr/>					
US 5494652	A	27-02-1996	DE	69215883 D1	23-01-1997
			DE	69215883 T2	26-06-1997
			WO	9306040 A1	01-04-1993
			EP	0605500 A1	13-07-1994
			FR	2681850 A1	02-04-1993