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(54) Titre : CATALYSEUR A BASE DE COBALT, ET SON UTILISATION DANS LE PROCEDE FISCHER-TROPSCH

(54) Title: CATALYST BASED ON COBALT AND ITS USE IN THE FISCHER-TROPSCH PROCESS

(57) Abrégé/Abstract:

A catalyst is described, which can be used in the Fischer-Tropsch process, essentially consisting of cobalt oxide supported on an inert carrier essentially consisting of alumina, characterized in that the above cobalt oxide essentially consists of crystals having an average size ranging from 20 to 80 Å. The preparation of the above catalyst and the Fischer-Tropsch process in its presence, are also described.

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Abstract

A catalyst is described, which can be used in the Fischer-Tropsch process, essentially consisting of cobalt oxide supported on an inert carrier essentially consisting of alumina, characterized in that the above cobalt oxide essentially consists of crystals having an average size ranging from 20 to 80 Å.

The preparation of the above catalyst and the Fischer-Tropsch process in its presence, are also described.

CATALYST BASED ON COBALT AND ITS USE IN THE FISCHER-TROPSCH PROCESS

The present invention relates to a catalyst based on 10 cobalt, its preparation and its use in the Fischer-Tropsch process.

The Fischer-Tropsch process is a process well-known to experts in the field, which essentially consists in the hydrogenation of CO to give hydrocarbons. The reaction conditions are also described in literature. 15

Catalysts which can be used in the Fischer-Tropsch process generally consist of metals of group VIII supported on a carrier, preferably selected from alumina, silica, titania and relative mixtures.

20 All research projects on the Fischer-Tropsch process are being increasingly orientated towards a greater selectivity to C_{9+} , particularly C_{22+} hydrocarbons, the latter also known as Fischer-Tropsch waxes.

Among these Fischer-Tropsch catalysts, Cobalt, which 25 is particularly effective in directing the reaction towards

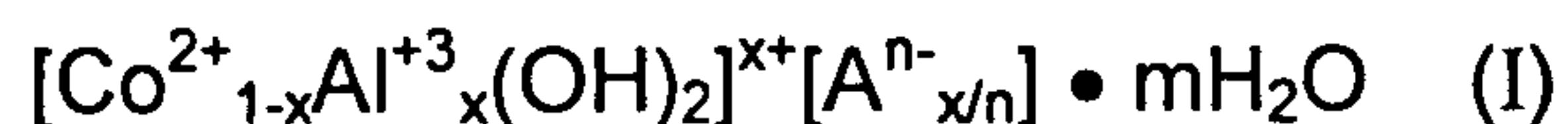
the formation of waxes, is becoming more and more widely used.

A particular catalyst based on cobalt supported on alumina has now been found, which is more selective, with respect to normal cobalt catalysts, towards the formation of waxes in the Fischer-Tropsch process.

In accordance with this, the present invention relates to a catalyst which can be used in the Fischer-Tropsch process, essentially consisting of cobalt oxide supported on an inert carrier essentially consisting of alumina, characterized in that the above cobalt oxide essentially consists of crystals having an average size ranging from 20 to 80 Å, preferably from 25 to 60 Å, even more preferably from 30 to 40 Å.

The invention is also directed to a catalyst used in a Fischer-Tropsch process, said catalyst comprising mixed oxide of cobalt and alumina, said cobalt oxide comprising crystals having an average size ranging from 20 to 80 Å, said catalyst being obtained by:

- 1) preparing an intermediate supported on alumina, having general formula (I)



wherein x ranges from 0.2 to 0.4, A is an anion, x/n is the number of anions necessary for neutralizing the positive charge, m ranges from 0 to 6; by dripping a solution of an Aluminium salt and a Cobalt salt, on a suspension of alumina, at a pH ranging from 6.6 to 7.2 and recovering the compound having general formula (I) supported on alumina thus obtained; and

- 2) calcinating the intermediate having general formula (I) at a temperature ranging from 300 to 500°C with the formation of crystalline cobalt oxide.

The above crystals of cobalt oxide can be optionally partly doped with Al atoms.

With respect to catalysts known in the art, the catalyst of the present invention has the characteristic of consisting of crystals with much lower dimensions, 20-80 Å for the catalyst of the present invention against the 120-180 Å of a catalyst prepared according to the conventional techniques. This allows a better dispersion of the cobalt on the carrier, with a consequent better contact, in reaction phase, between the catalyst and reagents.

In addition to cobalt oxide, the catalyst of the pres-

ent invention may optionally also contain, in a much lower quantity than the cobalt, metals normally known as promoters, such as Si, Zr, Ta, Zn, Sn, Mn, Ba, Ca, La, Ve, W. Promoters are used for improving the structural stability 5 of the carrier itself.

One or more activity promoters with a different effect on the catalytic performances as described in the art (see for example B. Jager, R. Espinoza in "Catalysis Today", 23, 1995, 21-22), can also be optionally present together with 10 the cobalt. For example, promoters such as K, Na, Mg, Sr, Cu, Mo, Ta, W and metals of group VIII essentially increase the activity. Ru, Zr, rare-earth oxides (REO), Ti increase the selectivity to high molecular weight hydrocarbons. Ru, REO, Re, Hf, Ce, U, Th favour the regenerability of cobalt 15 catalysts.

As far as the alumina is concerned, this can have any phase form selected from eta, gamma, delta, theta, alpha and relative mixtures, in the presence of or without one or more structural stability promoters selected from those described above. In the preferred embodiment, the alumina is 20 in γ or δ form, and relative mixtures.

The surface area of the alumina is that which is normal in catalytic carriers, i.e. from 20 to 300 m^2/g , preferably from 50 to 200 m^2/g (BET), whereas the average dimensions of the alumina itself range from 1 to 300 μm . 25

The cobalt content of the catalyst of the present invention ranges from 2 to 50% by weight, preferably from 5 to 20% by weight, 100 being the total weight of the carrier and cobalt (plus possible promoters). When present, the promoters are in a quantity not higher than 20% by weight with respect to the cobalt, preferably 10% by weight.

Before being used in the Fischer-Tropsch process, the catalyst of the present invention should be activated by means of the usual procedures, for example by reduction of cobalt oxide to metallic cobalt in the presence of hydrogen.

In accordance with this, the present invention relates to a process for obtaining a catalyst comprising mixed oxide of cobalt and alumina, said cobalt oxide comprising crystals having an average size ranging from 20 to 80 Å, said process comprising the following steps of:

- 1) preparing an intermediate supported on alumina, having general formula (I)



wherein x ranges from 0.2 to 0.4, A is an anion, x/n is the number of anions necessary for neutralizing the positive charge, m ranges from 0 to 6; by dripping a solution of an Aluminium salt and a Cobalt salt, on a suspension of alumina, at a pH ranging from 6.6 to 7.2 and recovering the compound having general formula (I) supported on alumina thus obtained; and

- 2) calcinating the intermediate having general formula (I) at a temperature ranging from 300 to 500°C with the formation of crystalline cobalt oxide.

With respect to the anion A^{n-} , this can be indifferently selected from inorganic anions (for example F^- , Cl^- , Br^- , I^- , ClO_4^- , NO_3^- , OH^- , IO_3^- , CO_3^{2-} , SO_4^{2-} , WO_4^{2-}), heteropolyacids (for example $PMo_{12}O_{40}^{3-}$, $PW_{12}O_{40}^{3-}$), organic acids (for example adipic, oxalic, succinic, malonic acid). In the preferred embodiment the anion A^{n-} is chosen from NO_3^- , OH^- , CO_3^{2-} . In an even more preferred embodiment A^{n-} is equal to CO_3^{2-} .

The compound having general formula (I) can be prepared according to various techniques known to experts in the field.

For example, the so-called precipitation technique can be used, according to which Co^{2+} and Al^{3+} are coprecipitated on alumina in the form of hydroxides. According to this technique, a solution of an Aluminum salt and a Cobalt salt, preferably an aqueous solution of the above salts, is dripped onto a suspension, preferably aqueous, of alumina. This operation must be effected maintaining the pH within a range of 6.6 to 7.2, preferably from 6.8 to 7.1, for example by the use of an aqueous solution of bicarbonate or soda. Alternatively, two separate solutions can be added; however, for the sake of simplicity, it is obviously preferable to use a single solution of the two salts. The compound having general formula (I) is recovered by means

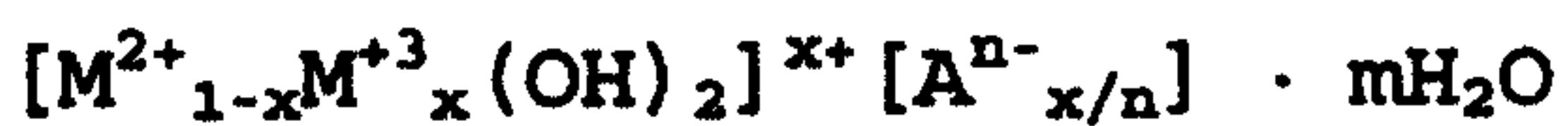
of filtration.

According to another less preferred embodiment, the so-called hydro-thermal technique can be used, which consists in treating freshly precipitated mixed cobalt and 5 aluminum hydroxides, or mechanical mixtures of the oxides, with water.

The compound having general formula (I) can be amorphous or crystalline. The ratio between the amorphous part and the crystalline part can be modified using known techniques 10 (for example by annealing). The crystalline part of the compound having general formula (I) has a structure similar to that of hydrotalcite.

Hydrotalcite is a mineral existing in nature and consists of an Al and Mg hydroxy-carbonate. A hydrotalcite-type 15 system has an analogous structure, but contains different elements.

The group of hydrotalcites can be represented by the following formula:



20 in the case of actual hydrotalcite $M^{2+} = Mg^{2+}$, $M^{3+} = Al^{3+}$ and $A^{n-} = CO_3^{2-}$.

One of the main characteristics of this group of compounds is the layered structure: layers of the brucite type, $Mg(OH)_2$, or $[M^{2+}_{1-x}M^{3+}_x(OH)_2]^{x+}$, in which a part of the 25 bivalent M^{2+} ions is substituted by trivalent M^{3+} ions, al-

ternate with anionic layers associated with a varying content of water ($A^{n-}_{x/n} \cdot mH_2O$). The anionic layers balance the positive charge of the hydroxide layers, the latter linked to the presence of trivalent ions.

5 In general, M^{2+} and M^{3+} can be ions of a varying nature, the only requisite is that they are able to insert themselves in the cavities left by the hydroxyls in a compact configuration of the brucite type (more simply, they must have an ionic radius similar to that of Mg^{2+}). The 10 value of x in the structural formula ranges from 0.2 to 0.4, preferably from 0.25 to 0.35. Outside this range pure hydroxides or other compounds with a different structure can be obtained. Formulations of the hydrotalcite type have the following ratios $M^{2+}/M^{3+} = 2$ and $M^{2+}/M^{3+} = 3$, as limits 15 in the composition, which do not imply differences in the lattice and structural parameters (only the internal distribution of the cations in the brucite layer changes). More information on compounds of the hydrotalcite type are contained in the review "Hydrotalcite-type anionic clays: 20 preparation, properties and applications" (Catalysis Today, 11(1991) 173-301).

Once the compound having general formula (I) has been prepared and isolated, before beginning step (2), it is preferable to subject the compound itself to a drying step 25 in order to reduce the quantity of water (or solvent) ad-

sorbed.

Step (2) consists in the calcination of the compound having general formula (I) at a temperature ranging from 300 to 500°C, preferably from 350 to 450°C.

5 As already mentioned, before being used in the Fischer-Tropsch process, the catalyst is subjected to reduction. This operation is carried out, in the preferred embodiment, by means of treatment with hydrogen, optionally diluted with inert gases, for example nitrogen. The reduction step is preferably effected at a temperature ranging from 300°C to 500°C, even more preferably from 320°C to 10 450°C. It is possible to operate either under pressure or at atmospheric pressure, the latter condition being preferred. The duration of the reduction process varies in relation to the experimental conditions (temperature, pressure, dilution or non-dilution of the hydrogen).
15

A further object of the present invention relates to a process for the preparation of prevalently C_{22+} hydrocarbons, or so-called Fischer-Tropsch waxes, characterized in 20 that it is carried out in the presence of the catalyst according to claim 1.

The Fischer-Tropsch process is the well-known reaction between CO and H₂, optionally diluted with CO₂ and/or N₂, to give prevalently C_{22+} hydrocarbons.

25 The reaction conditions are described in literature.

For example, the temperatures can range from 170°C to 400°C, preferably from 180°C to 250°C, whereas the pressures can vary from 1 to 100 bars, preferably from 15 to 40 bars. The CO/H₂ ratio can vary from 0.5/1 to 4/1, preferably from 1.87/1 to 2.5/1, the stoichiometric value (more or less 3%) being preferred.

The catalyst of the present invention can be used either in a fixed bed reactor or in a slurry-type reactor.

The following examples are provided for a better understanding of the present invention.

EXAMPLES

The examples considered refer to two catalysts based on cobalt called A (comparative catalyst) and B. For both the carrier is γ - δ alumina Condea Scca 5-170.* The morphological characteristics of the carrier are indicated below:

Table 1

Alumina carrier (Al ₂ O ₃)	
Crystalline phase	60% γ - 40% δ
Surface area (m ² /g)	162
Pore specific volume (cm ³ /g)	0.46
Particle-size distribution	50% < 82 Å (d _p)

PREPARATION OF COMPARATIVE CATALYST A

Catalyst A is prepared by means of the incipient wetness technique.

* Trademark

100 g of γ - δ alumina Condea Scca 5-170 are impregnated with 100 cc of an aqueous solution of $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (0.3 moles). The impregnated carrier is left to dry in an oven at 80°C for 16 h, and is then subjected to calcination in 5 muffle in a stream of air according to the following thermal profile:

- from 25 to 350°C (10°C/min)
- isotherm at 350°C (30 min)
- from 350 to 400°C (5°C/min)
- 10 - isotherm 400°C (240 min)
- from 400 to 25°C (10°C/min).

The dried catalyst is characterized by means of XRD analysis.

The calcined catalyst is characterized by means of 15 elemental analysis, XRD analysis and XPS analysis.

CHARACTERIZATION OF THE DRIED SAMPLE A

The dried sample A is subjected to XRD analysis. The spectrum reveals the presence of large crystals of hexavalent cobalt nitrate used during the impregnation phase as 20 cobalt precursor, as well as the γ , δ phase of the Condea Scca 5-170* alumina carrier.

CHARACTERIZATION OF THE CALCINED SAMPLE A

The calcined catalyst A is subjected to elemental analysis in order to determine the actual cobalt content 25 introduced with the preparation. The cobalt content deter-

* Trademark

mined corresponds exactly to 14.2% by weight.

The sample is also subjected to XRD analysis and the spectrum reveals the presence, in addition to the γ - δ phase of the carrier, of a spinel phase which can be attributed 5 to the species $\text{Co}(\text{Co}_{2-x},\text{Al}_x)\text{O}_4$.

The Al % contained is defined on the basis of a structural refinement of the spectrum. This value refers to the % of cationic sites of the pure spinel phase of Cobalt oxide (Co_3O_4) occupied by aluminum ions and indicates interactions between oxide precursor and carrier. Table 2 specifies the data obtained from the XRD spectrum relating to weight % of the spinel phase, % of Al ions present in the spinel phase, and dimensions of the crystallites. The latter was determined using the Scherrer method.

15 Table 2. XRD analysis of catalyst A

Spinel phase %	Al ion %	Crystal size (\AA)
20.3	7.2	135

The calcined catalyst A is finally subjected to XPS analysis. This spectroscopic technique can provide information as to the composition and morphology of the system, 20 with respect to the surface layers of the sample.

Table 3 indicates the results of the analysis:

Table 3. XPS analysis of catalyst A

B.E. Co 2p _{3/2} (eV)	ΔE Co (2p _{1/2} - 2p _{3/2}) (eV)	Co 2p / Al 2s
780	15.1	0.15

B.E. = binding energy

The first two columns of Table 3 identify the presence of Co_3O_4 on the surface of the catalyst whose composition is well defined by XRD analysis; the third column refers to 5 the ratio between the intensity of the XPS signal relating to cobalt 2p and the intensity of the signal relating to aluminum 2s of the carrier and gives important indications regarding the distribution of the cobalt precursor with respect to the carrier. The higher the value of this ratio, 10 the more uniformly distributed the phase containing cobalt will be on the surface of the carrier. In the case of catalyst A the ratio is equal to 0.15, which means that the cobalt oxide is not uniformly distributed on the surface of the carrier, but rather can induce the assumption of the 15 presence of islands of cobalt oxide distributed at random on the surface. This is a characteristic common to alumina supported systems prepared according to the traditional techniques, often confirmed by SEM images.

The cobalt reduction percentage, determined analytically 20 on the sample reduced at 400°C for 16 h in a stream of pure H_2 , corresponds to 35%.

PREPARATION OF CATALYST B

Unlike catalyst A, catalyst B of the present invention is prepared according to the synthesis method which produces 25 the compound having general formula (I), having a hy-

drotalcite-type structure, and its subsequent calcination.

Two aqueous solutions are prepared: one of $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ (molar ratio Co/Al 3:1) having a concentration equal to 1 Mol ($\text{M}^{\text{n}+}$)/liter and one of NaHCO_3 having a concentration equal to 1.3 Moles/liter.

The two solutions are slowly and contemporaneously dripped at room temperature into a single container, into which the γ - δ alumina Condea Scca 5-170* carrier has been dispersed (55 g in about 200 ml of distilled water). The objective is to obtain, after calcination, a cobalt content ranging from 13 to 15% by weight. During the addition of the two solutions, the pH, controlled by means of a glass electrode, remains constant, between 6.8 and 7.1.

The solid is recovered by filtration on a būchner and repeatedly and abundantly washed with distilled water; it is then left to dry in an oven at 80°C for 16 hours. Calcination is subsequently effected in muffle in a stream of air.

The calcination treatment is carried out under the following conditions:

- from 25 to 350°C ($10^{\circ}\text{C}/\text{min}$)
- isotherm at 350°C (30 min)
- from 350 to 400°C ($5^{\circ}\text{C}/\text{min}$)
- isotherm 400°C (240 min)
- from 400 to 25°C ($10^{\circ}\text{C}/\text{min}$).

* Trademark

Analogously to example A, the dried catalyst B is characterized by means of XRD analysis and the calcined catalyst B is characterized by means of elemental analysis, XRD analysis and XPS analysis.

5 CHARACTERIZATION OF THE DRIED CATALYST B

After drying, catalyst B is subjected to XRD analysis. The spectrum reveals the presence of a hydrotalcite-type crystalline phase, well formed in spite of the relatively low crystal size. This phase can be represented 10 by the formula:



The results of the XRD characterization are indicated in Table 4.

The quantity of hydrotalcite phase calculated corresponds to a lower cobalt content than the total quantity 15 determined via elemental analysis, and it should therefore be assumed that the missing part is otherwise located in the aluminum carrier, forming an undetectable fraction, well dispersed in the carrier.

20 It cannot be excluded that there is still a hydrotalcite phase present in the form of crystallites of such small dimensions as to be invisible to XRD analysis.

In any case it should be noted that the presence of cobalt nitrate hexahydrate is not observed, contrary to 25 what is confirmed in the dried reference catalyst.

CHARACTERIZATION OF THE CALCINED CATALYST B

The calcined catalyst B is subjected to elemental analysis in order to determine the actual content of cobalt introduced with the preparation. The cobalt content determined exactly corresponds to 14.7% by weight.

The sample is also subjected to XRD analysis and the spectrum reveals the presence of two spinel phases, one is γ -Al₂O₃ deriving from the carrier, the other is a mixed Co/Al/O oxide. The results are indicated in Table 4 below.

10 Table 4. XRD analysis of catalyst B

B	Phase	a (Å)	V (Å ³)	Crystal size (Å)
Dried	Hydrotalcite	3.074	188.3	65
Calcined	Co/Al/O	8.086	528.7	45

15 Unlike the dried sample, the quantity of mixed oxide is congruent with the probable cobalt content.

The two phases have cell parameters extremely close to those of the respective references (Table 5).

Table 5. Comparison of the cell parameters with literature data.

	a (Å)	V (Å ³)
Co ₃ O ₄ lit. ref.	8.084	528.3
γ -Al ₂ O ₃ lit. ref.	7.924	497.5
Co ₂ AlO ₄ lit. ref.	8.086	528.7
Phase γ -Al ₂ O ₃ cat. B	7.920	496.8
Phase Co/Al/O cat. B	8.086	528.7

25

Calculation on the basis of three possible formulae leads to the results indicated in Table 6.

Table 6

Hypothetical phase type	Calculated Co %
Co ₃ O ₄	17.6
Co ₂ AlO ₄	13.5
CoAl ₂ O ₄	8

The most probable formulation, also on the basis of the elemental datum, is Co₂AlO₄, perhaps a little more enriched with cobalt, better represented by the formula 10 Co_{2+x}Al_{1-x}O₄.

It can be concluded that with calcination, the cobalt dispersed in the carrier, not visible to XRD, contributes, together with the visible hydrotalcite, to the formation of the mixed Co/Al/O oxide.

15 The size of the crystallites of the Co/Al/O spinel phase present in catalyst B is considerably lower than that registered in catalyst A for the Co(Co_(2-x), Al_x)O₄ spinel.

Catalyst B, after calcination, is subjected to XPS analysis analogously to the procedure for catalyst A. The 20 relative results are indicated in Table 7:

Table 7. XPS analysis of catalyst B

B.E. Co 2p _{3/2} (eV)	ΔE Co (2p _{1/2} - 2p _{3/2}) (eV)	Co 2p / Al 2s
781	15.2	1.73

B.E. = binding energy

25 As for catalyst A, the first two columns indicate the

presence of the compound Co_3O_4 on the surface of the carrier; the B.E. value however referring to Co 2p equal to 781 eV is much higher than that observed for catalyst A, which can be explained by the presence of an "aluminated" compound (confirmed and better described as a mixed Co/Al/O system by XRD analysis). The most interesting aspect is linked to the $\text{Co}2\text{p}/\text{Al}2\text{s}$ ratio (third column) which is unusually high, 10 time higher than that registered for catalyst A. Such a high value can only be explained by a homogeneous distribution of the cobalt precursor on the surface of the carrier. The precursor covers the carrier hiding the XPS signal.

The cobalt reduction percentage, analytically determined on the sample reduced at 400°C for 16 h in a stream 15 of pure H_2 , corresponds to 40%.

DESCRIPTION OF THE CATALYTIC TESTS

Catalysts A and B were used in a reactivity test in an FTR reactor fed in continuous with a mixture of CO and H_2 under the conditions indicated in Table 8 below.

20 Table 8. Operating conditions of the catalytic tests

Reaction temperature:	200-240 °C
Total pressure:	21 abs. bars
H_2/CO ratio:	2/1
Space velocity:	1.3 – 3.0 $\text{NI} / \text{cc}_{\text{cat}}/\text{h}$

The catalyst is charged, in the pre-established quantities, into the fixed bed tubular reactor. The activation of the catalyst is carried out *in situ* by means of hydrogen reduction (2 Nl/h/1cat) and nitrogen (1N1/h/1cat) at a temperature ranging from 320-450°C and a pressure of 1 bar for 16 hours. At the end the reactor is cooled in a stream of nitrogen.

During this phase the system is brought to the final operating pressure of 22 abs. bars. The reagent mixture, consisting of H₂ and CO in a stoichiometric ratio of 2:1, is introduced by the progressive feeding of CO-H₂ and reduction of the feeding of N₂ as indicated in Table 9.

Table 9. Feeding conditions in activation phase

Time range (hours)	CO flow rate (Nl/h)	H ₂ flow rate (Nl/h)	N ₂ flow rate (Nl/h)
0 - 0.5	10	20	200
0 - 0.5	10	20	150
0 - 0.5	10	20	100
0 - 0.5	10	20	50
0 - 0.5	10	20	0

At the end of the activation phase, the system proves to be totally without gaseous diluent (nitrogen) and under the conditions of pressure, space velocity, H₂/CO ratio indicated in the table. The temperature is subsequently

raised to the reaction temperature in about 15 hours. The waxes produced are collected in suitable containers at a temperature of 110°C. The effluent gas from the reactor passes through a meter and a subsequent sampling system for 5 gas-chromatographic analysis. The solid and liquid effluents are analyzed with appropriate gas-chromatographic equipment for the total quantification. In order to normalize the catalytic activity data of the various tests, with respect to the actual cobalt content, the yield to products 10 containing carbon (hydrocarbons and CO₂) normalized for the actual moles of cobalt present in the catalyst and for the time units: defined as Co-TY (Cobalt-Time Yield) = moles of converted CO/moles of total Co/hour, is used as a comparative parameter.

15 **EXAMPLE 1. Comparative catalyst A**

The catalytic activity tests are indicated in Table 10.

Table 10. Performances of catalyst A

Catalyst amount (cc)	20	
T.O.S. (h)	396	730
GHSV (NI / l _{cat} / h)	1500	1500
Temperature (°C)	206	215
Test pressure (abs. bar)	21	21
actual H ₂ /CO	2.01	2.01
Co % conversion	33.3	51.8
Co-TY (mol conv. CO / h / mol Co)	4.1	6.5
C ₂₊ productivity (C ₂₊ / h / Kg _{cat})	109.8	168.2
C ₉₊ productivity (C ₉₊ / h / Kg _{cat})	41.7	119.5
C ₂₂₊ productivity (C ₂₂₊ / h / Kg _{cat})	23.3	53.1

EXAMPLE 2. Catalyst B

The data of the catalytic activity test are indicated in the following table and compared, with iso-conversion, with the data relating to example 1 to evaluate the positive effect, in terms of productivity to higher hydrocarbons (C_{22+}), of the catalyst of the present invention.

Table 11. Performances of catalyst B

Catalyst amount (cc)	20	
T.O.S. (h)	139	187
GHSV (NI / l_{cat} / h)	1500	1500
Temperature (°C)	225	235
Test pressure (abs. bar)	21	21
actual H_2/CO	2.08	2.08
Co % conversion	33.3	46.6
Co-TY (mol conv. CO / h / mol Co)	4.2	5.8
C_{2+} productivity (C_{2+} / h / Kg_{cat})	123.3	179.7
C_{9+} productivity (C_{9+} / h / Kg_{cat})	96.0	136.3
C_{22+} productivity (C_{22+} / h / Kg_{cat})	50.2	66.2

SPECIFIC ACTIVITY

Table 12 indicates the specific activity values of the two catalysts in terms of TOF (turnover frequency), or converted CO moles per active site in the time unit.

Table 12. Specific activity

catalyst	T (°C)	TOF $\times 10^2$ s ⁻¹
A	222	4.0
B	219	5.8

CATALYTIC PERFORMANCES**COMMENT ON TABLES 10, 11, 12**

The use of the catalyst based on cobalt supported on alumina of the present invention allows higher catalytic performances to be obtained, in terms of selectivity to heavy products, than those obtained with a catalyst obtained by the deposition of cobalt on the same alumina carrier, according to the traditional incipient wetness technique.

Example 2 shows how catalyst B, prepared according to the method of the present invention, allow higher productivity values, in terms of C_{9+} and C_{22+} , to be obtained (Table 11) with respect to those obtained with comparative catalyst A (Example 1, Table 10).

The comparison between the performances of catalysts A and B is effected at the same CO conversion level equal to 30 and 50%.

In conclusion, catalyst B of the present invention favours the formation of hydrocarbon products with a higher molecular weight, thus forming an additional advantage for transformation processes of Fischer-Tropsch synthesis products.

In order to evaluate the catalytic activity, the specific reactivity of each case intended as TOF, "turnover frequency", must be considered.

A comparison between the TOF data (Table 12) of homologous catalysts, i.e. A and B (the same carrier), demonstrates how the B system, prepared according to the method of the present invention, has a greater specific activity.

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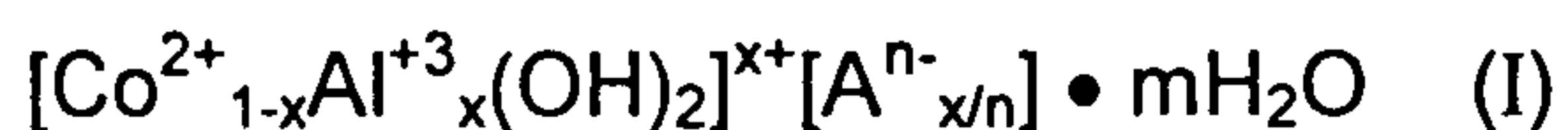
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CLAIMS

1. A catalyst used in a Fischer-Tropsch process, said catalyst comprising mixed oxide of cobalt and alumina, said cobalt oxide comprising crystals having an average size ranging from 20 to 80 Å, said catalyst being obtained by:

- 1) preparing an intermediate supported on alumina, having general formula (I)

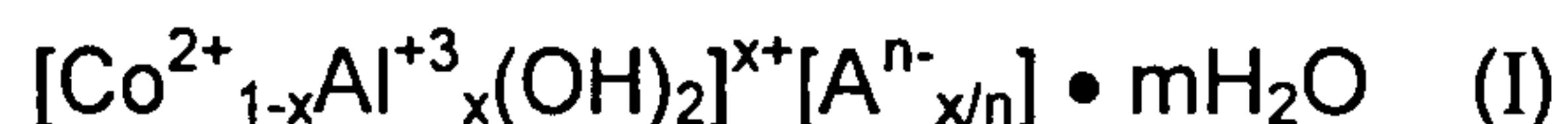


wherein x ranges from 0.2 to 0.4, A is an anion, x/n is the number of anions necessary for neutralizing the positive charge, and m ranges from 0 to 6; by dripping a solution of an Aluminium salt and a Cobalt salt, on a suspension of alumina, at a pH ranging from 6.6 to 7.2 and recovering the compound having general formula (I) supported on alumina thus obtained; and

- 2) calcinating the intermediate having general formula (I) at a temperature ranging from 300 to 500°C with the formation of crystalline cobalt oxide.
2. The catalyst according to claim 1, wherein the cobalt oxide crystals have an average size of 25 to 60 Å.
3. The catalyst according to claim 2, wherein the cobalt oxide crystals have an average size of 30 to 40 Å.
4. The catalyst according to any one of claims 1 to 3, wherein the alumina is in γ or δ form, and relative mixtures.

5. The catalyst according to any one of claims 1 to 4, wherein the cobalt content ranges from 2 to 50% by weight.
6. The catalyst according to claim 5, wherein the cobalt content ranges from 2 to 20% by weight.
7. A process for obtaining a catalyst comprising mixed oxide of cobalt and alumina, said cobalt oxide comprising crystals having an average size ranging from 20 to 80 Å, said process comprising the following steps of:

- 1) preparing an intermediate supported on alumina, having general formula (I)



wherein x ranges from 0.2 to 0.4, A is an anion, x/n is the number of anions necessary for neutralizing the positive charge, and m ranges from 0 to 6; by dripping a solution of an Aluminium salt and a Cobalt salt, on a suspension of alumina, at a pH ranging from 6.6 to 7.2 and recovering the compound having general formula (I) supported on alumina thus obtained; and

- 2) calcinating the intermediate having general formula (I) at a temperature ranging from 300 to 500°C with the formation of crystalline cobalt oxide.
8. The process according to claim 7, wherein x ranges from 0.25 to 0.35.
9. The process according to claim 7, wherein m = 4.

10. The process according to claim 7, characterized in that $A = CO_3$.
11. The process according to claim 7, wherein step (2) is preceded by a drying step of the compound having general formula (I).
12. The process according to any one of claims 7 to 11, wherein the pH ranges from 6.8 to 7.1.
13. The process according to any one of claims 7 to 12, characterized in that the calcination is effected at a temperature ranging from 320 to 450°C.
14. A process for the production of hydrocarbon waxes according to the Fischer-Tropsch process, said process being characterized in that it is carried out in the presence of a catalyst according to any one of claims 1 to 6.