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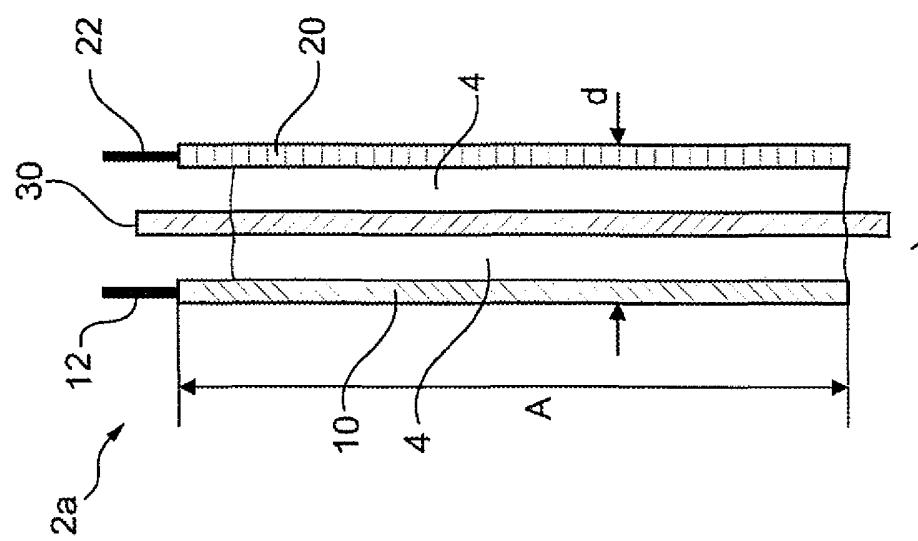
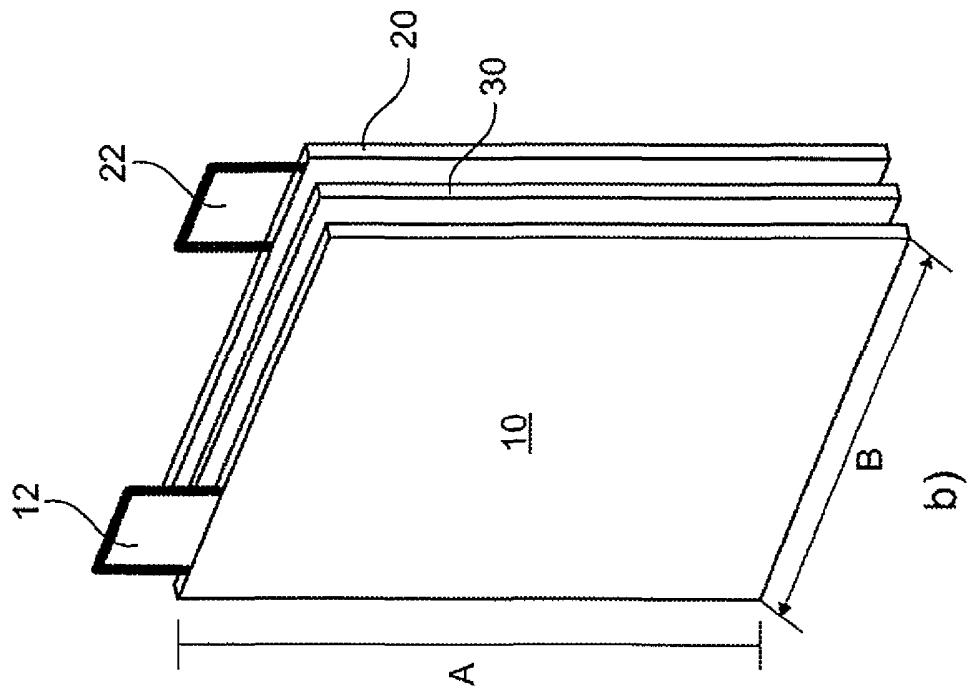
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(54) Title of the Invention: **Electrolyte for a battery**
Abstract Title: **Lithium Cell Electrolyte**

(57) The present disclosure relates to an electrolyte for an electrochemical cell and an electrochemical cell comprising such an electrolyte. The electrolyte comprises at least one conductive salt comprising lithium ions, at least one solvent and at least one wetting agent e.g. a fluoropolymer or ionic surfactant or flurosurfactant. The electrochemical cell comprises at least one anode, at least one cathode and at least one separator arranged between the at least one anode and the at least one cathode. The electrolyte may be filled between the at least one anode and the at least one cathode

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a)
Fig. 1

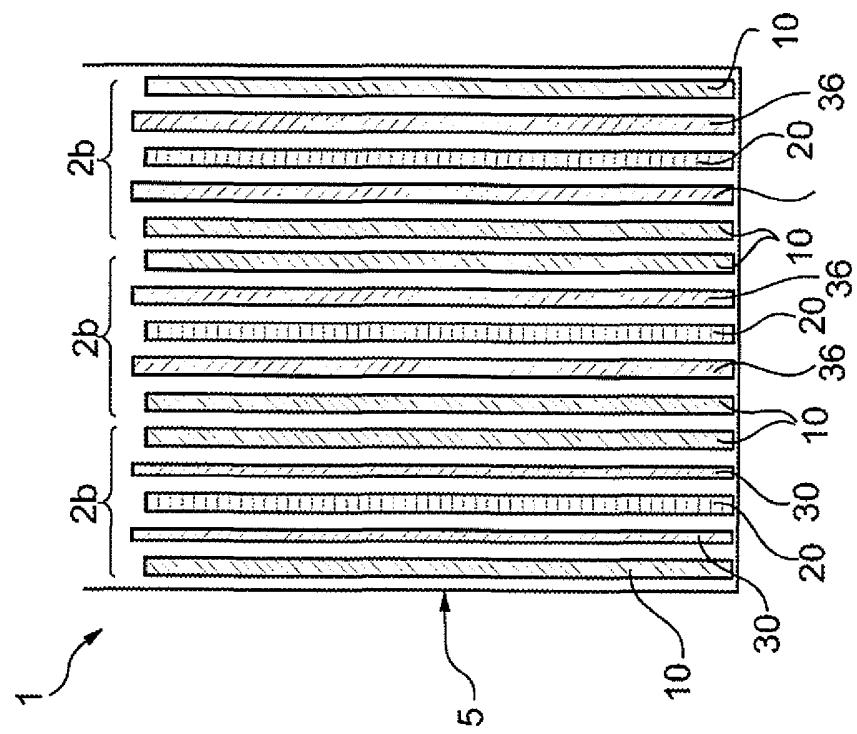


Fig. 2b

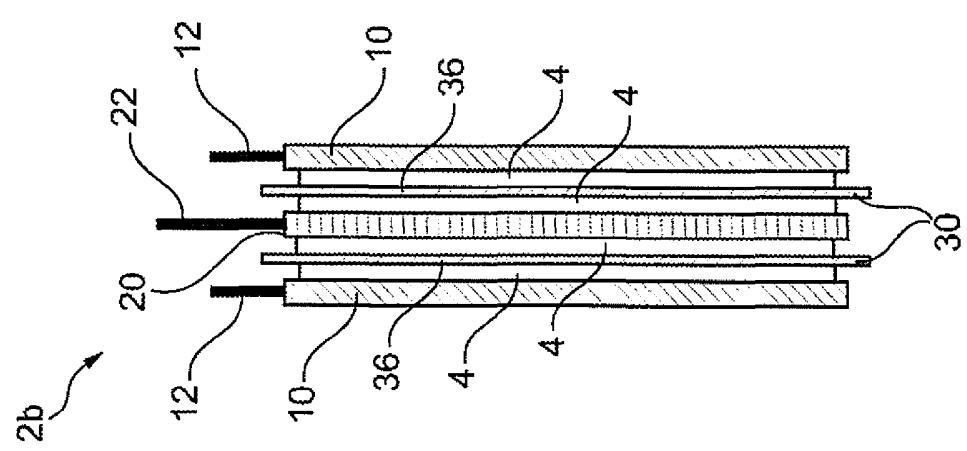


Fig. 2a

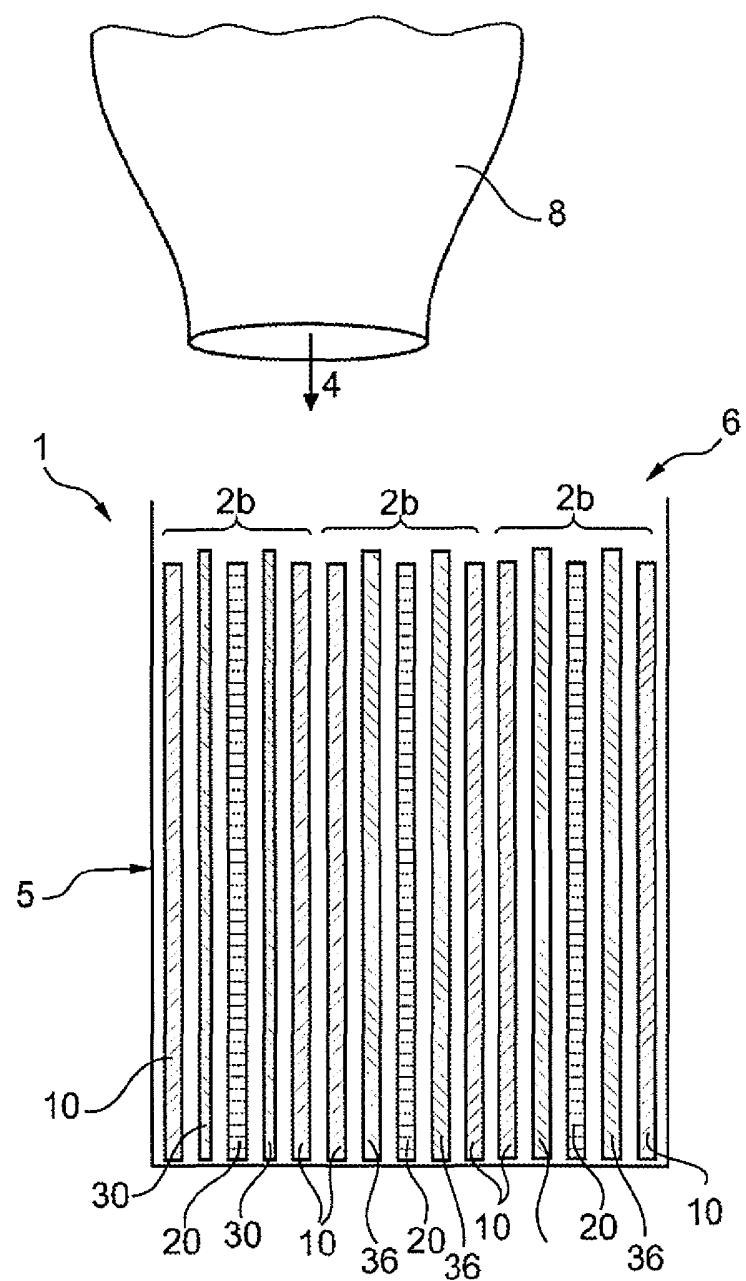


Fig. 3

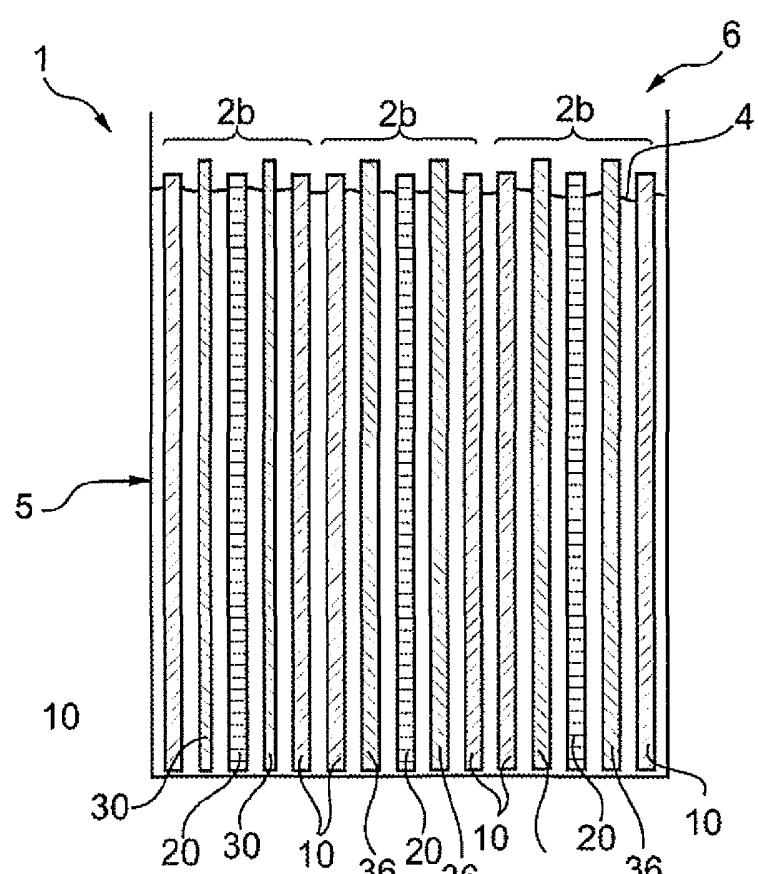


Fig. 4

Description

5 [0001] The present disclosure relates to rechargeable lithium-ion containing electrochemical cells and batteries and to a manufacturing method thereof. In particular, the present disclosure relates to an electrolyte for a large format electrochemical cell and to a method for filling the electrolyte into the electrochemical cell used in rechargeable lithium-ion containing batteries.

10 [0002] Lithium-ion containing rechargeable batteries, also called lithium ion secondary batteries or lithium ion batteries, are advantageous because of their large capacities, their extended life times, the absence of a memory effect and have been widely used for small sized applications. Lithium containing rechargeable batteries are widely used for many applications, and have shown to be particularly useful in mobile phones, mobile computers and other

15 electronic devices.

[0003] The use of the lithium-ion containing rechargeable batteries, however, is today limited to smaller cells with limited capacities. Just few large scale lithium cell batteries have been presented on the market so far, although there is an increasing need for large scale and

20 high capacity lithium batteries, for example for the use in electric vehicles or as energy buffers or storages in green energy power plants, such as solar farms or wind farms. Storage of large amounts of electricity is becoming an increasing need for future energy solutions.

[0004] However, manufacturing of large format lithium containing batteries has not been

25 possible in a manner that allow cost effective mass production of the large format lithium cells. Present production processes are very time consuming mainly due to the time consuming filling of the electrolyte into the cell and do not allow a cost-effective production of large format lithium containing electrochemical cells or batteries.

30 [0005] WO 02/091497 describes non-ionic surfactants as additives to the electrolyte in lithium ion batteries. These additives are used for improving impedance properties of the

battery, but the disclosed method does not accelerate the filling of electrolyte into an electrochemical cell.

[0006] It is an object of the present invention to improve the manufacturing of such

5 electrochemical cells.

Summary of the Invention

[0007] The present disclosure relates to an electrolyte for an electrochemical cell and an

10 electrochemical cell comprising such an electrolyte. The electrolyte comprises at least one conductive salt comprising lithium ions, at least one solvent and at least one wetting agent. The electrochemical cell comprises at least one anode, at least one cathode and at least one separator arranged between the at least one anode and the at least one cathode. The electrolyte may be filled between the at least one anode and the at least one cathode.

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[0008] The invention also relates to a method for manufacturing an electrochemical cell.

The method comprises providing at least one anode, at least one cathode and at least one separator between the at least one anode and the at least one cathode, and filling an electrolyte between the anode and the cathode, wherein the electrolyte comprises at least one wetting

20 agent.

[0009] Using the wetting agent in the electrolyte allows faster filling of the electrochemical

cell. Using the wetting agent in the electrolyte enables filling of large format electrochemical

25 cells, even with small distances between the anode and the cathode. The amount of time

necessary for filling the electrolyte in the electrochemical cell between the at least one anode and the at least one cathode is considerably reduced. The use of the wetting agent in the electrolyte allows a homogenous distribution of electrolyte between the at least one anode and the at least one cathode, in particular without gas bubbles or other inhomogeneities.

30

[0010] A large format electrochemical cell may have at least one dimension of about 100 mm or more. For example, at least one of a cathode, an anode and a separator between the anode and the cathode may have at least one dimension of about 100 mm or more, for example a surface area of about 0.01 m² or more. The present invention makes the manufacture of much larger electrochemical cells possible.

[0011] The at least one anode and the at least one cathode of the electrochemical cell may be arranged at a distance of about 1 mm or less, in particular 0.5 mm or less. The at least one anode and/or the at least one cathode may have a thickness of about 100 μm or less, for example 50 μm or less, thus allowing the manufacture of space and material reduced electrochemical cells with high capacities.

[0012] The conductive salt comprising lithium ions may be or may comprise at least one of LiPF₆, LiClO₄, , LiBF₄, LiAsF₆ and LiPF₃(CF₂CF₃), Lithium bis [1,2-oxalato(2-)-O,O'] borate (LiBOB) based electrolytes, Lithium tris (pentafluoroethyl) trifluorophosphate Li[(C₂F₅)₃PF₃] short LiFAP, LiF₄C₂O₄, LiFOP, LiPF₄(C₂O₄), LiF₄OP, LiCF₃ SO₃ , LiC₄F₉SO₃, Li(CF₃SO₂)₂N, Li(C₂ F₅ SO₂)₂ N, LiSCN and LiSbF₆, Lithium-trifluormethansulfonat ("Li-Triflat"), Lithiumimide (Lithium-bis (perfluoralkylsulfonyl)-imide) sowie Lithiummethide (Lithium-tris (perfluoralkylsulfonyl) methide), LiIm(BF₃)₂, high voltage LiTDI, LiPDI and LiHDI (lithium salts of 2-perfluoroalkylo-4,5-dicyanoimidazole), LiAlO₄, LiAlCl₄, LiCl and LiI and the like.

[0013] The at least one wetting agent may be or may comprise a fluoropolymer. Possible examples for fluoropolymers comprise commercially available perflourinated alkyl ethoxylates such as Zonyl SFO, Zonyl SFN und Zonyl SF300 (E. I. DuPont). Li- thium-3-[(1H,1H,2H,2H-fluoralkyl)thio]-propionat, Zonyl FSA ©, Du Pont). Other fluoropolymers that may be used with the present disclosure comprise semi-fluorinated acryl polymer EGC-1700, Fluoromethacrylate, long-chain perfluoroacrylates, tetrafluorethylene, hexafluoropropylene, silane-coupling agent with perfluoropolyether (PFPE-S), (perfluoroalkyl)ethyl methacrylate-containing acrylic polymers, butyl methacrylate-*co*-perfluoroalkyl acrylate, semifluorinated fluorocarbon diblock copolymer poly(butyl methacrylate-*co*-perfluoroalkyl acrylate), n-perfluorononane, perfluoropropyleneoxyde, polytetrafluoroethylene, poly(tetrafluoroethylene-*co*-hexafluoropropylene), perfluorobutyl (PFB), perfluoromethyl, perfluoroethyl or a combination thereof.

30

[0014] The at least one wetting agent may be or comprise an ionic surfactant, in particular an anionic surfactant, such as a fluorosurfactant. Commercially available examples of fluorosurfactants that may be used with the present disclosure comprise, but are not limited to, fluorosurfactants distributed by DuPont under the product name Zonyl SFK, Zonyl SF-62 or

distributed by 3M Company under the product name FLURAD FC 170, FC 123, or L-18699A.

[0015] Other commercially available products that may be used as fluorosurfactant comprise

5 3M Company products distributed under the product name Novec F-C4300, 3M FC-4430, 3M FC-4432, or 3M FC-4434.

[0016] The at least one wetting agent may be provided in the electrolyte at a final concentration of about 5000 ppm (parts per million) or less, in particular in a concentration of 10 about 500 ppm or less to limit foam formation. At least one wetting agent may be provided in the electrolyte at a final concentration of about 5 ppm or more, in particular of about 50 ppm or more. These concentrations have been found give good results with respect to fast and homogenous filling of the electrolyte into a pre-assembled cell.

15 [0017] The solvent may be a non-aqueous solvent. The non-aqueous solvent may comprise any combination of ionic liquids. The non-aqueous solvent may comprise at least one of a cyclic carbonate, a cyclic ester, a linear carbonate, ether or a combination thereof. The non-

aqueous solvent may be an organic solvent comprising at least one solvent selected from the group consisting of propylene carbonate (PC), ethylene carbonate (EC), diethyl carbonate (DEC), dimethyl carbonate (DMC), dipropyl carbonate (DPC), dimethyl sulfoxide, acetonitrile, dimethoxyethane, diethoxyethane, tetrahydrofuran, N-methyl-2-pyrrolidone (NMP), ethylmethyl carbonate (EMC), methylpropyl carbonate (MPC), fluoroethylene carbonate (FEC), γ -butyrolactone (GBL), methyl formate, ethyl formate, propyl formate, methyl acetate, ethyl acetate, propyl acetate, pentyl acetate, methyl propionate, ethyl

20 propionate, propyl propionate, butyl propionate, 1,2-butylene carbonate, 2,3-butylene carbonate, 1,2-pentylene carbonate and 2,3-pentylene carbonate or a combination thereof.

Brief Description of the Figures

30 [0018] The following description gives examples of embodiments of the present disclosure and is made with respect to the attached Figures in a purely exemplifying and non limiting manner, wherein:

[0019] Figures 1a and 1b show an example of an electrochemical cell;

[0020] Figure 2a shows a second example of an electrochemical cell and Fig 2b shows how electrochemical cells can be stacked to form a battery;

5 [0021] Figure 3 shows filling a battery comprising a plurality of stacked electrochemical cells with electrolyte;

[0022] Figure 4 shows the filled battery.

10 Detailed Description

[0023] Figure 1 shows an example of an electrochemical cell 2 that can be used with the present disclosure. The electrochemical cell 2 comprises two electrodes, an anode 10, and a cathode 20. The anode 10 and the cathode 20 are separated by a separator 30. The anode 10 and the cathode 20 may be made from any material known in the art of electrochemical cells. For example, the anode 10 may comprise a collector and a carbon or graphite coating or lithium titanate oxide or any other lithium metal alloys, but the anode is not limited to such materials. The collector may be made from copper, aluminium, stainless steel, titanium or any other material known in the art. The cathode 20 may comprise a cathode-collector made from aluminium, stainless steel, titanium or any other material known in the art and may comprise a metal oxide layer such as aluminium oxide or other materials known in the art such as lithium cobalt oxide or other metals oxides, but not limited to such materials.

[0024] The anode 10 and the cathode 20 have electrical contacts 12, 22 for electrically contacting the respective electrode.

[0025] The separator 30 may be a ceramic separator as known in the art. The invention is, however, not limited to the above materials and any electrode or separator material known, such as for example polyolefin-based or polyester-based materials can be used with the present disclosure.

[0026] The electrochemical cell 2a may be a large format electrochemical cell. An electrochemical cell may be called a large format electrochemical cell if at least one of the electrodes 10, 20 and the separator 30 between the electrodes have a length A and/or a width

B of at least about 10 cm or more. For example the length A and the width B of the electrodes 10, 20 can be about 10 to about 20 cm. The length A may be different than the width B allowing rectangular shapes or any other shape desired. The shape of the electrode may be adapted to the application of the electrochemical cell or battery and may be adapted to a 5 particular casing.

[0027] In the shown example, the distance D between the anode 10 and the cathode 20 is less than 1 mm. For example, the distance between an anode collector of the anode 10 and a cathode connector of the cathode 20 may about 400 μm or less.

10

[0028] Each one of the electrodes 10, 20 of the anode 10 and the cathode 20 may be made of a foil material of a thickness of about less than 50 μm . In particular the foils may have a thickness of about 10 to 20 μm . For example, an aluminium foil may be used for the cathode 20 and a copper foil may be used for the anode 10.

15

[0029] The electrochemical cell 2a is filled with an electrolyte 4 that is in contact with the anode 10 and the cathode 20.

20

[0030] Fig. 2a shows an electrochemical cell 2b that differs from the electrochemical cell 2a in that at both sides of the cathode 20 a separator 30 and an anode 10 are arranged. The electrolyte 4 is inserted between each anode 10 and the cathode 20. This allows closer stacking of the electrochemical cells 2b in a battery 1 and requires less cathode material. The electrical contacts 12, 22 are omitted in the figures for clarity reasons.

25

[0031] A plurality of the electrochemical cells 2a as shown in Figure 1a and 1b or a plurality of electrochemical cells 2b as shown in Figure 2a may be stacked on top of each other to form a rechargeable battery 1. Fig. 2b illustrates how a plurality of electrochemical cells 2b can be stacked in a housing, pack or pouch 5. The number of electrochemical cells 2 stacked can be varied according to the application of the rechargeable battery 1. In the example show, three 30 electrochemical cells 2b are shown for illustrative purposes stacked to form a rechargeable battery 2, but the number of electrochemical cells 2a, 2b can be much higher. For example, a battery 2 may comprise up to about 500 electrochemical cells 2a, 2b.

[0032] The electrochemical cells 2a as shown in Figure 1A and 1B may simply be stacked on top of each other and the electrodes 10, 20 may be separated from each other using a separator material.

5 [0033] However, other stacking methods are also possible and applicable with the present invention. Figures 2-4 show electrochemical cells 2b in bicell-configuration. The cell can also be implemented in monocell-configuration, bipolar-configuration, as wound or Z-stacked cell. The active masses or active materials can be coated single-sided or double-sided to the collector. Other stacking methods may be applied as well, such as alternating stacking of
10 anodes and cathodes, each with a separator material in between. By doing this, it is possible to use both surfaces of the anode and of the cathode.

[0034] Fig. 2b shows a plurality of electrochemical cells 2b stacked in a package or pouch 5 in bicell-configuration, prior to filling electrolyte into the electrochemical cells 2b.

15

[0035] Fig. 3 shows how the electrolyte 4 may be inserted in the electrochemical cells 2a, 2b. The electrochemical cells 2a, 2b may be packed in a pouch 5 that is closed on all sites except the top side 6 using a dosing apparatus 8 such as a needle or the like. Fig. 3 shows a bicell-configuration of three pairs of electrochemical cells 2b, wherein the contacts 12, 22 are
20 omitted for clarity reasons. The dosing apparatus 8 allows inserting an pre-determined amount of electrolyte 4 into the electrochemical cells 2a,2b. Inserting the electrolyte 4 in the electrochemical cells 2a, 2b packed in the pouch 5 may be performed under vacuum conditions, for example at a pressure of about 10 to 500 mbar abs. The electrolyte 4 may be injected from one side only, substantially simplifying the injection procedure.

25

[0036] It is important to have a very homogenous distribution of electrolyte 4 between the anode 10 and the cathode 20, in particular, no bubbles or other errors shall be present between the anode 10 and the cathode 20, as this will lead to undesired defects and less battery capacities. The electrolyte 4 used in lithium containing batteries 1 may comprise a non-aqueous solvent such as, for example, a cyclic carbonate, a cyclic ester, a linear carbonate, an ether, or a combination thereof. Other organic solvents may be used.
30

[0037] The electrolyte 4 for lithium ion batteries 1 also comprises conductive lithium salts such as for example LiClO₄, LiPF₆, LiBF₄, LiAsF₆ and LiPF₃(CF₂CF₃), Lithium bis [1,2-

oxalato(2-)O,O'] borate (LiBOB) based electrolytes, LiF₄C₂O₄, LiFOP, LiPF₄(C₂O₄), LiF₄OP, LiCF₃ SO₃, LiC₄F₉SO₃, Li(CF₃SO₂)₂N, Li(C₂F₅SO₂)₂N, LiSCN and LiSbF₆, LiAlO₄, LiAlCl₄, LiCl and LiI or a combination thereof. Other known lithium salts may be used as well.

5

[0038] The electrolyte 4 comprises a wetting agent. The wetting agent is used to homogenously wet the surfaces of the anodes 10, the cathodes 20 and the separator 30 and to obtain a homogeneous distribution of electrolyte 4 inside the electrochemical cells 2a, 2b. The wetting agent may be or may comprise a fluoropolymer, in particular a fluorosurfactant.

10 Possible examples for fluoropolymers comprise commercially available perflourinated alkyl ethoxylates such as Zonyl SFO, Zonyl SFN und Zonyl SF300 (E. I. DuPont). Li-thium-3-[(1H,1H,2H,2H-fluoralkyl)thio]-propionat, Zonyl FSA ©, Du Pont).

15 [0039] Commercially available examples of fluorosurfactants that may be used with the present disclosure comprise but are not limited to fluorosurfactants distributed by DuPont under the product name Zonyl SFK, Zonyl SF-62 or distributed by 3M Company under the product name FLURAD FC 170, FC 123, or L-18699A. Other commercially available product that may be used as fluorosurfactant comprise 3M Company products distributed under the product name Novec F-C4300, 3M FC-4430, 3M FC-4432, or 3M FC-4434.

20

[0040] Other wetting agents that may be used with the present disclosure comprise semi-fluorinated acryl polymer EGC-1700, Fluoromethacrylate, long-chain perfluoroacrylates, tetrafluorethylene, hexafluoropropylene, silane-coupling agent with perfluoropolyether (PFPE-S), (perfluoroalkyl)ethyl methacrylate-containing acrylic polymers, butyl methacrylate-*co*-perfluoroalkyl acrylate, semifluorinated fluorocarbon diblock copolymer poly(butyl methacrylate-*co*-perfluoroalkyl acrylate), n-perfluorononane, perfluoropropyleneoxyde, polytetrafluoroethylene, poly(tetrafluoroethylene-*co*-hexafluoropropylene), perfluorobutyl (PFB), perfluoromethyl, perfluoroethyl or a combination thereof.

25

[0041] All of the above wetting agents may be used alone or in any combination.

30 [0042] The wetting agents, fluoropolymers or fluorosurfactants may be used at a concentration of about 5 ppm (parts per million) to about 5000 ppm.

[0043] The use of the wetting agent in the electrolyte results in an even and homogeneous distribution of the electrolyte 4 in the electrochemical cell 2a, 2b. The use of the wetting agent allows reducing the filling times considerably and allows to manufacture large format lithium ion batteries in acceptable time scales suitable for mass production.

[0044] Figure 4 shows a sealed battery pack 1, wherein the opening 6 of the pouch 5 has been closed after filling the battery pack 1 with electrolyte 4 has been completed.

10 [0045] It is obvious to a person skilled in the art that other possibilities than pouches 5 exist to pack the electrochemical cells 2a, 2b. For example a battery housing from known plastics materials may be used.

15 [0046] It is obvious to a person skilled in the art that a plurality of battery packs 1 may be combined to increase the capacity and/or voltage of the battery.

[0047] The electrolyte of the present disclosure may be used with any type of electrochemical cells and a person skilled in the art may adapt the properties of the electrolyte to different applications, i.e. to the size and material of the electrochemical cells used.

20

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Claims

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1. An electrolyte (4) for an electrochemical cell (2a, 2b), the electrolyte (4) comprising:
 - at least one conductive salt comprising lithium ions,
 - at least one solvent and
 - at least one wetting agent.

10

2. The electrolyte (4) of claim 1, wherein the at least one wetting agent comprises a fluoropolymer.
3. The electrolyte (4) of claim 1 or 2, wherein the at least one wetting agent comprises an ionic surfactant.
4. The electrolyte (4) of any one of the preceding claims, wherein the at least one wetting agent comprises a fluorosurfactant.
5. The electrolyte (4) of any one of the preceding claims, wherein a concentration of the at least one wetting agent in the electrolyte is about 5000 ppm or less.
6. The electrolyte (4) of any one of the preceding claims, wherein a concentration of the at least one wetting agent in the electrolyte is about 5 ppm or more.
7. The electrolyte (4) of any one of the preceding claims, wherein the solvent is a non-aqueous solvent.
8. The electrolyte (4) of any one of the preceding claims, wherein the solvent comprises at least one of a cyclic carbonate, a cyclic ester, a linear carbonate, an ether or a combination thereof and/or any combination of ionic liquids.

25

30

9. An electrochemical cell (2a, 2b) comprising an electrolyte according to any one of the preceding claims.

10. The electrochemical cell (2a, 2b) of claim 9, wherein an anode-collector of an anode (10) and a cathode-collector of a cathode (20) of the electrochemical cell (2a, 2b) are arranged at a distance of about 1 mm or less.

11. The electrochemical cell (2a, 2b) of claim 9 or 10, wherein at least one of an anode (10) or a cathode (20) has a surface area of about 0.01 m² or more.

10

12. The electrochemical cell (2a, 2b) of any one of claims 9 to 11, wherein at least one of an anode or a cathode has a thickness of about 300 µm or less.

15

13. The electrochemical cell (2a, 2b) of any one of claims 9 to 12, wherein at least an anode, a cathode and a separator are laminated to each other.

20

14. A method for manufacturing an electrochemical cell (2a, 2b), the method comprising the steps of:

- providing at least one anode (10), at least one cathode (20) and at least one separator (30) between the at least one anode (10) and the at least one cathode (20); and
- filling an electrolyte (4) between the anode (10) and the cathode (20), wherein the electrolyte (4) comprises at least one wetting agent.

25

15. The method of claim 14, wherein the step of filling the electrolyte (4) between the anode (10) and the cathode (20) comprises injecting the electrolyte (4) from one side of the at least one anode (10), the at least one cathode (20) and the at least one separator (30).

30

16. The method of claim 15, further comprising placing the at least one anode (10), the at least one cathode (20) and the at least one separator (30) in a pouch (5) with one open side and wherein injecting the electrolyte (4) comprises injection the electrolyte (4) through the open side of the pouch (5).

17. The method of any one of claims 14 to 16, wherein filing the electrolyte between the anode and the cathode is performed under vacuum.
18. The method of any one of claims 14 to 17, wherein the providing the at least one anode, the at least one cathode and the at least one separator comprise laminating the at least one anode, the at least one cathode and the at least one separator to each other.
5
19. The method of any one of claims 14 to 18, wherein the electrolyte is an electrolyte according to any one of claims 1 to 8.

10

Application No: GB1013977.2

Examiner: Tony Martin

Claims searched: 1-13

Date of search: 29 September 2010

Patents Act 1977: Search Report under Section 17

Documents considered to be relevant:

Category	Relevant to claims	Identity of document and passage or figure of particular relevance
X	1 at least	US2010/0047695 A Smart see claims 1 and 2
X	" "	WO2009/035085 A Daikin see abstract
X	" "	WO2010/004012 A Giroud see abstract

Categories:

X	Document indicating lack of novelty or inventive step	A	Document indicating technological background and/or state of the art.
Y	Document indicating lack of inventive step if combined with one or more other documents of same category.	P	Document published on or after the declared priority date but before the filing date of this invention.
&	Member of the same patent family	E	Patent document published on or after, but with priority date earlier than, the filing date of this application.

Field of Search:

 Search of GB, EP, WO & US patent documents classified in the following areas of the UKC^X :

Worldwide search of patent documents classified in the following areas of the IPC

H01M

The following online and other databases have been used in the preparation of this search report

On line databases WPI,EPODOC

International Classification:

Subclass	Subgroup	Valid From
H01M	0010/0567	01/01/2010