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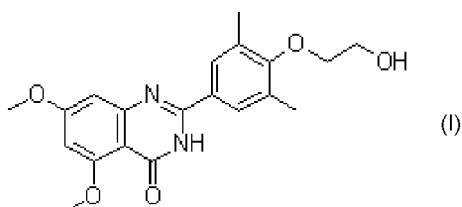
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(54) Title: CRYSTAL FORM OF BROMODOMAIN PROTEIN INHIBITOR DRUG, PREPARATION METHOD AND USE THEREOF

(54) 发明名称: 一种溴结构域蛋白抑制剂药物的晶型及其制备方法和用途



(57) Abstract: Provided are a crystal form of a bromodomain protein inhibitor 2-[4-(2-hydroxyethoxy)-3,5-dimethylphenyl]-5,7-dimethoxyquinazolin-4(3H)-one represented by Formula I, a preparation method and a use thereof. The present invention belongs to the field of medicine. The crystal forms are crystal form CS2, crystal form CS8, crystal form CS13, crystal form CS20, crystal form CS1, crystal form CS7, crystal form CS9, crystal form CS11, and crystal form CS4, and can be used to prepare drugs for treating cardiovascular, cholesterol, or lipid-related disorders.

(57) 摘要: 一种式 I 所示的溴结构域蛋白抑制剂 2-[4-(2-羟基乙氧基)-3,5-二甲基苯基]-5,7-二甲氧基喹唑啉-4(3H)-酮的晶型及其制备方法和用途, 属于医药领域。所述晶型为晶型 CS2、晶型 CS8、晶型 CS13、晶型 CS20、晶型 CS1、晶型 CS7、晶型 CS9、晶型 CS11 及晶型 CS4, 可用于制备治疗心血管、胆固醇或脂质相关紊乱疾病的药物。

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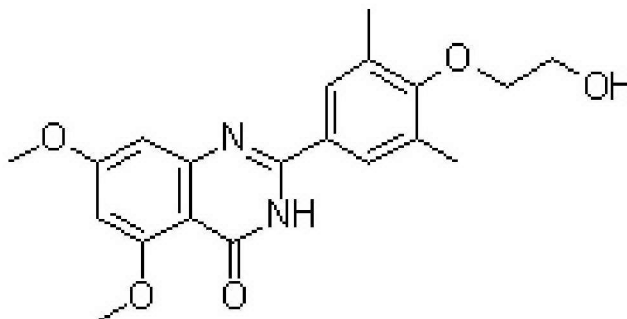
CRYSTALLINE FORMS OF A BROMODOMAIN AND EXTRATERMINAL PROTEIN INHIBITOR DRUG, PROCESSES FOR PREPARATION THEREOF, AND USE THEREOF

TECHNICAL FIELD

The present disclosure relates to the field of pharmaceutical chemistry, particularly relates to novel crystalline forms of 2-[4-(2-hydroxyethoxy)-3,5-dimethylphenyl]-5,7-dimethoxyquinazolin-4(3H)-one, processes for preparation and use thereof.

BACKGROUND

2-[4-(2-hydroxyethoxy)-3,5-dimethylphenyl]-5,7-dimethoxyquinazolin-4(3H)-one, also known as apabetalone, is a bromodomain and extraterminal (BET) protein inhibitor developed by Resverlogix in Canada. Apabetalone can be used in the treatment of cardiovascular, cholesterol or lipid-related disorders. Especially in the treatment of atherosclerosis, acute coronary syndrome and predecessor diabetes, apabetalone has shown significant curative effect. The structure of apabetalone is shown as follows:



Cardiovascular diseases, also known as circulatory system diseases, can be subdivided into acute diseases and chronic diseases and generally relate to atherosclerosis. Cardiovascular disease is one of the most serious diseases threatening human life worldwide. Its morbidity and mortality rate have surpassed that of tumour diseases and ranked first. In China, there are about 290 million patients suffering from cardiovascular diseases, and the number of patients is increasing year by year. However, people's need for drugs of cardiovascular diseases have not been fulfilled, and new drugs still need to be developed continuously. Studies have shown that apabetalone can inhibit BRD4 of BET family, thereby regulating the expression of apolipoprotein A-1 (ApoA-1) and the synthesis of high density lipoprotein cholesterol, and realizing the treatment of cardiovascular related diseases. No patent or literature about apabetalone crystalline forms is

found after searching patents and literatures published locally and abroad.

Different crystalline forms of solid chemical drugs have different solubility and stability, and can affect drug's in vivo dissolution and absorption, which will further affect drug's clinical efficacy. The inventors of the present disclosure surprisingly discovered crystalline form CS2, form CS8, form CS13, form CS20, form CS1, form CS7, form CS9, form CS11 and form CS4 after conducting a large number of experiments, which provides a new and better choice for the preparation of pharmaceutical preparations containing apabetalone and is of great significance for drug development.

The discussion of documents, acts, materials, devices, articles and the like is included in this specification solely for the purpose of providing a context for the present invention. It is not suggested or represented that any or all of these matters formed part of the prior art base or were common general knowledge in the field relevant to the present invention as it existed before the priority date of each claim of this application.

Where the terms "comprise", "comprises", "comprised" or "comprising" are used in this specification (including the claims) they are to be interpreted as specifying the presence of the stated features, integers, steps or components, but not precluding the presence of one or more other features, integers, steps or components, or groups thereof.

SUMMARY

Form CS2, Form CS8, Form CS13, Form CS20, Form CS1, Form CS7, Form CS9, Form CS11 and Form CS4 were discovered by the inventors of the present disclosure after a lot of experiments and research. The above polymorphs have good stability. The polymorphs are stable for at least two weeks under the conditions of 25 °C/60%RH and/or 40 °C/75%RH, preferably at least 6 weeks, and more preferably at least 10 months. The novel polymorphs discovered by the inventors of the present disclosure have advantages in simple and repeatable preparation method, high purity, good solubility and low hygroscopicity, which makes the polymorphs meet the requirements of medical use and is suitable for production and application.

A main aspect of the present disclosure is to provide novel crystalline forms of apabetalone, processes for preparation and use thereof.

According to an aspect of the present disclosure, crystalline form CS2 is provided (hereinafter referred to as "Form CS2"). Said Form CS2 is a hydrate.

The X-ray powder diffraction pattern of Form CS2 shows characteristic peaks at 2theta values of $11.5^{\circ} \pm 0.2^{\circ}$, $6.6^{\circ} \pm 0.2^{\circ}$ and $8.8^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS2 shows one or two characteristic peaks at 2theta values of $5.1^{\circ} \pm 0.2^{\circ}$ and $15.3^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS2 shows characteristic peaks at 2theta values of $5.1^{\circ} \pm 0.2^{\circ}$ and $15.3^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS2 shows characteristic peaks at 2theta values of $11.5^{\circ} \pm 0.2^{\circ}$, $6.6^{\circ} \pm 0.2^{\circ}$, $8.8^{\circ} \pm 0.2^{\circ}$, $5.1^{\circ} \pm 0.2^{\circ}$, $15.3^{\circ} \pm 0.2^{\circ}$, $13.3^{\circ} \pm 0.2^{\circ}$, $20.2^{\circ} \pm 0.2^{\circ}$, $23.1^{\circ} \pm 0.2^{\circ}$ and $25.3^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Without any limitation being implied, in a specific embodiment, the X-ray powder diffraction pattern of Form CS2 is substantially as depicted in FIG 1A.

According to another aspect of the present disclosure, a process for preparing Form CS2 of apabetalone is also provided. The process comprises: Adding apabetalone solid into alcohols, a mixture of alcohols and ethers, a mixture of alcohols and ketones or a mixture of alcohols and arenes. Heating to dissolve the solid, filtering and cooling the solution to obtain a solid. The obtained solid is Form CS2 of apabetalone.

Furthermore, said alcohol includes methanol; said ether includes 2-methyltetrahydrofuran; said arene includes toluene; said ketone includes methyl isobutyl ketone; said heating temperature is $50\text{-}100^{\circ}\text{C}$; said cooling temperature is $-20\text{-}5^{\circ}\text{C}$.

Form CS2 of the present disclosure has following advantages:

1) At present, there is no patent or literature about crystalline forms of apabetalone. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS2, which is suitable for drug development.

2) Form CS2 of the present disclosure has good solubility in SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids). Especially in SGF, the solubility of the Form CS2 is as high as 0.61 mg/mL at 24 hours. Higher solubility is beneficial to improve drug's blood concentration and bioavailability, which is of great significance for drug research.

According to another aspect of the present disclosure, crystalline form CS8 of apabetalone is provided (hereinafter referred to as Form CS8). Said Form CS8 is an anhydrate.

The X-ray powder diffraction pattern of Form CS8 shows characteristic peaks at 2theta values of $13.5^{\circ} \pm 0.2^{\circ}$, $7.8^{\circ} \pm 0.2^{\circ}$, $22.5^{\circ} \pm 0.2^{\circ}$ and $11.4^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS8 shows one or two characteristic

peaks at 2theta values of $25.9^{\circ} \pm 0.2^{\circ}$ and $13.1^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS8 shows characteristic peaks at 2theta values of $25.9^{\circ} \pm 0.2^{\circ}$ and $13.1^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS8 shows one or two characteristic peaks at 2theta values of $28.1^{\circ} \pm 0.2^{\circ}$ and $20.2^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS8 shows characteristic peaks at 2theta values of $28.1^{\circ} \pm 0.2^{\circ}$ and $20.2^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS8 shows characteristic peaks at 2theta values of $23.9^{\circ} \pm 0.2^{\circ}$, $13.5^{\circ} \pm 0.2^{\circ}$, $7.8^{\circ} \pm 0.2^{\circ}$, $22.5^{\circ} \pm 0.2^{\circ}$, $11.4^{\circ} \pm 0.2^{\circ}$, $25.9^{\circ} \pm 0.2^{\circ}$, $13.1^{\circ} \pm 0.2^{\circ}$, $28.1^{\circ} \pm 0.2^{\circ}$, $20.2^{\circ} \pm 0.2^{\circ}$ and $9.7^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction pattern of Form CS8 is substantially as depicted in FIG 2A.

According to another aspect of the present disclosure, a process for preparing Form CS8 of apabetalone is also provided. The process comprises: Adding apabetalone solid into a single solvent of halohydrocarbons or a mixture of solvents of halohydrocarbons and alcohols. Heating to dissolve the solid, filtering and cooling the solution to obtain a solid. The obtained solid is Form CS8 of apabetalone.

Furthermore, said halohydrocarbon includes dichloromethane; said alcohol includes isopropanol; said volume ratio of halohydrocarbon and isopropanol is preferably 4:1; said heating temperature is $40\text{--}60^{\circ}\text{C}$, preferably 50°C ; said cooling temperature is $-20\text{--}5^{\circ}\text{C}$.

Form CS8 of the present disclosure has following advantages:

1) At present, there is no patent or literature about crystalline forms of apabetalone. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS8, which is suitable for drug development.

2) Form CS8 of the present disclosure has good long-term stability. Form CS8 is stable for at least 2 weeks when stored under the condition of $25^{\circ}\text{C}/60\%\text{RH}$, which is conducive to long-term storage. Form CS8 has excellent stability, which can ensure the quality of the drug will not be affected as polymorphic transition will not occur in the process of preparation, transportation and storage. It is of great significance to ensure the efficacy and safety of the drug and prevent the occurrence of adverse drug reactions.

3) Form CS8 of the present disclosure has low hygroscopicity. The weight gain of Form CS8 of the present disclosure at $80\%\text{RH}$ is 0.34% . Form CS8 is slightly hygroscopic. Due to the low hygroscopicity, it is not necessary to control the environmental humidity in preparation processes,

and no strict requirement is needed for packaging and storage, which is beneficial to long-term storage and industrial production of drugs and reduces the cost. Due to the unstrict requirement on the storage conditions, the cost of storage and quality control will be greatly reduced, which has strong economic value and is more suitable for medicinal use.

5 According to another aspect of the present disclosure, crystalline form CS13 is provided (hereinafter referred to as "Form CS13"). Said Form CS13 is a hydrate.

The X-ray powder diffraction pattern of Form CS13 shows characteristic peaks at 2theta values of $5.1^{\circ}\pm 0.2^{\circ}$, $12.5^{\circ}\pm 0.2^{\circ}$ and $17.1^{\circ}\pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS13 shows one or two or three
0 characteristic peaks at 2theta values of $6.4^{\circ}\pm 0.2^{\circ}$, $8.5^{\circ}\pm 0.2^{\circ}$ and $25.7^{\circ}\pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS13 shows characteristic peaks at 2theta values of $6.4^{\circ}\pm 0.2^{\circ}$, $8.5^{\circ}\pm 0.2^{\circ}$ and $25.7^{\circ}\pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS13 shows one or two characteristic
5 peaks at 2theta values of $7.8^{\circ}\pm 0.2^{\circ}$ and $16.0^{\circ}\pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS13 shows characteristic peaks at 2theta values of $7.8^{\circ}\pm 0.2^{\circ}$ and $16.0^{\circ}\pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS13 shows characteristic peaks at 2theta values of $5.1^{\circ}\pm 0.2^{\circ}$, $12.5^{\circ}\pm 0.2^{\circ}$, $17.1^{\circ}\pm 0.2^{\circ}$, $6.4^{\circ}\pm 0.2^{\circ}$, $8.5^{\circ}\pm 0.2^{\circ}$, $25.7^{\circ}\pm 0.2^{\circ}$, $7.8^{\circ}\pm 0.2^{\circ}$ and $16.0^{\circ}\pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction
20 pattern of Form CS13 is substantially as depicted in FIG 3A.

According to another aspect of the present disclosure, a process for preparing Form CS13 of apabetalone is also provided. The process comprises: Dissolving apabetalone solid in a mixture of ethers and water or a mixture of ketones and water. Filtering the solution and then evaporating the filtrate at room temperature to obtain a solid. The obtained solid is Form CS13 of apabetalone.

25 Furthermore, said ether includes tetrahydrofuran; said ketone includes acetone; said volume ratio of ether and water is preferably 4:1; said volume ratio of ketone and water is preferably 9:1.

Form CS13 of the present disclosure has following advantages:

1) At present, there is no patent or literature about crystalline forms of apabetalone. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found
30 Form CS13, which is suitable for drug development.

2) Form CS13 of the present disclosure has good solubility in SGF and FeSSIF. Highly soluble

crystalline forms are conducive to improving drug's plasma concentration and bioavailability.

According to another aspect of the present disclosure, crystalline form CS20 of apabetalone is provided (hereinafter referred to as Form CS20). Said Form CS20 is an acetic acid solvate.

The X-ray powder diffraction pattern of Form CS20 shows characteristic peaks at 2theta values of $8.4^{\circ} \pm 0.2^{\circ}$, $18.9^{\circ} \pm 0.2^{\circ}$ and $13.5^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS20 shows one or two or three characteristic peaks at 2theta values of $11.3^{\circ} \pm 0.2^{\circ}$, $9.4^{\circ} \pm 0.2^{\circ}$ and $5.6^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS20 shows characteristic peaks at 2theta values of $11.3^{\circ} \pm 0.2^{\circ}$, $9.4^{\circ} \pm 0.2^{\circ}$ and $5.6^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS20 shows one or more characteristic peaks at 2theta values of $26.3^{\circ} \pm 0.2^{\circ}$, $20.1^{\circ} \pm 0.2^{\circ}$, $20.6^{\circ} \pm 0.2^{\circ}$ and $24.4^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS20 shows characteristic peaks at 2theta values of $26.3^{\circ} \pm 0.2^{\circ}$, $20.1^{\circ} \pm 0.2^{\circ}$, $20.6^{\circ} \pm 0.2^{\circ}$ and $24.4^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS20 shows one or two or three characteristic peaks at 2theta values of $14.5^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$ and $22.8^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS20 shows characteristic peaks at 2theta values of $14.5^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$ and $22.8^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS20 shows characteristic peaks at 2theta values of $8.4^{\circ} \pm 0.2^{\circ}$, $18.9^{\circ} \pm 0.2^{\circ}$, $13.5^{\circ} \pm 0.2^{\circ}$, $11.3^{\circ} \pm 0.2^{\circ}$, $9.4^{\circ} \pm 0.2^{\circ}$, $5.6^{\circ} \pm 0.2^{\circ}$, $26.3^{\circ} \pm 0.2^{\circ}$, $20.1^{\circ} \pm 0.2^{\circ}$, $20.6^{\circ} \pm 0.2^{\circ}$, $24.4^{\circ} \pm 0.2^{\circ}$, $14.5^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$ and $22.8^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction pattern of Form CS20 is substantially as depicted in FIG 4A.

According to another aspect of the present disclosure, a process for preparing Form CS20 of apabetalone is also provided. The process comprises: Dissolving apabetalone solid in a mixture of acetic acid and nitriles or a mixture of acetic acid and esters. Filtering the solution and then evaporating the filtrate at room temperature to obtain a solid. The obtained solid is Form CS20 of apabetalone.

Furthermore, said nitrile includes acetonitrile; said ester includes ethyl acetate; said volume ratio of nitrile and acetic acid is 9:1; said volume ratio of ester and acetic acid is 4:1.

Form CS20 of the present disclosure has following advantages:

1) At present, there is no patent or literature about crystalline forms of apabetalone. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS20, which is suitable for drug development.

2) Form CS20 has good long-term stability. Form CS20 is stable for at least 2 weeks when stored under the conditions of 25 °C/60% RH and 40 °C/75% RH. Form CS20 has excellent stability, which can ensure that the quality of the drug will be affected as polymorphic transition will not occur in the process of preparation, transportation and storage. It is of great significance to ensure the efficacy and safety of the drug and prevent the occurrence of adverse drug reactions.

3) Form CS20 of the present disclosure has good solubility in SGF and FeSSIF. Highly soluble crystalline forms are conducive to improving drug's blood concentration and bioavailability.

According to another aspect of the present disclosure, crystalline form CS1 of apabetalone is provided (hereinafter referred to as Form CS1). Said Form CS1 is an anhydrate.

The X-ray powder diffraction pattern of Form CS1 shows characteristic peaks at 2theta values of $6.1^{\circ} \pm 0.2^{\circ}$, $12.3^{\circ} \pm 0.2^{\circ}$, $26.1^{\circ} \pm 0.2^{\circ}$ and $26.8^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS1 shows one or two or three characteristic peaks at 2theta values of $16.4^{\circ} \pm 0.2^{\circ}$, $18.5^{\circ} \pm 0.2^{\circ}$ and $23.2^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS1 shows characteristic peaks at 2theta values of $16.4^{\circ} \pm 0.2^{\circ}$, $18.5^{\circ} \pm 0.2^{\circ}$ and $23.2^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS1 shows one or more characteristic peaks at 2theta values of $13.0^{\circ} \pm 0.2^{\circ}$, $14.1^{\circ} \pm 0.2^{\circ}$, $17.1^{\circ} \pm 0.2^{\circ}$ and $24.5^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS1 shows characteristic peaks at 2theta values of $13.0^{\circ} \pm 0.2^{\circ}$, $14.1^{\circ} \pm 0.2^{\circ}$, $17.1^{\circ} \pm 0.2^{\circ}$ and $24.5^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS1 shows characteristic peaks at 2theta values of $6.1^{\circ} \pm 0.2^{\circ}$, $12.3^{\circ} \pm 0.2^{\circ}$, $26.1^{\circ} \pm 0.2^{\circ}$, $26.8^{\circ} \pm 0.2^{\circ}$, $16.4^{\circ} \pm 0.2^{\circ}$, $18.5^{\circ} \pm 0.2^{\circ}$, $23.2^{\circ} \pm 0.2^{\circ}$, $13.0^{\circ} \pm 0.2^{\circ}$, $14.1^{\circ} \pm 0.2^{\circ}$, $17.1^{\circ} \pm 0.2^{\circ}$, $24.5^{\circ} \pm 0.2^{\circ}$ and $20.5^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction pattern of Form CS1 is substantially as depicted in FIG 5A.

According to another aspect of the present disclosure, a process for preparing Form CS1 of apabetalone is also provided. The process comprises:

Method 1: Dissolving step: Dissolving apabetalone solid in a solvent to obtain a clear solution.

Precipitation step: Adding the prepared solution to an anti-solvent or adding an anti-solvent to the prepared solution to obtain a solid. The obtained solid is Form CS1 of apabetalone.

Furthermore, said solvent is a single solvent or a mixture of solvents selected from the group consisting of tetrahydrofuran, chloroform, dimethyl sulfoxide and dimethyl acetamide; said anti-solvent is a single solvent or a mixture of solvents selected from the group consisting of n-heptane, methyl tert-butyl ether, toluene, water and acetonitrile.

Method 2: Dissolving step: Dissolving apabetalone solid in solvent at 40-60 °C to obtain a clear solution.

Precipitation step: Cooling the prepared solution to -20-5 °C to obtain a solid. The obtained solid is Form CS1 of apabetalone.

Furthermore, said solvent is tetrahydrofuran, acetone, a mixture of tetrahydrofuran and methyl tert-butyl ether, a mixture of ethyl acetate and acetone, a mixture of acetonitrile and N, N-dimethyl formamide.

Preferably, said volume ratio of tetrahydrofuran and methyl tert-butyl ether is 2:1; said volume ratio of ethyl acetate and acetone is 1:1; said volume ratio of acetonitrile and N, N-dimethyl formamide is 9:1.

Form CS1 of the present disclosure has following advantages:

1) At present, there is no patent or literature about apabetalone's crystalline forms. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS1, which is suitable for drug development.

2) Form CS1 has good long-term stability and mechanical stability. Form CS1 is stable for at least 10 months when stored under the conditions of 25 °C/60% RH and 40 °C/75% RH. The stability of crystalline form is very important for drug development. Form CS1 has excellent stability, which can ensure that the quality of the drug will not be affected as the polymorphic transition will not occur in the process of preparation, transportation and storage. It is of great significance to ensure the efficacy and safety of the drug and prevent the occurrence of adverse drug reactions. In addition, Form CS1 doesn't change after manual grinding, which indicates that Form CS1 has good mechanical stability. The grinding of APIs is usually needed in production process. Good grinding stability can reduce the risk of crystallinity change and polymorphic transition of APIs during production process.

3) Form CS1 of the present disclosure is almost non hygroscopic. The weight gain of Form CS1

of the present disclosure at 80% RH is 0.13% and the crystalline form doesn't change after DVS test. For almost non-hygroscopic crystalline form, it is not necessary to control the environmental humidity in production process. There is no special strict requirements for packaging and storage conditions, which saves costs and is suitable for industrial production and long-term storage of drugs. Due to the unstrict requirement on the storage conditions, the cost of material storage and quality control will be greatly reduced, which has strong economic value and is more suitable for medicinal use.

4) Form CS1 of the present disclosure has a good dissolution rate. In drug development, Form CS1 with fast dissolution rate can accelerate drug's in vivo dissolution. By adjusting the excipients, it is possible to control the rapid action of drugs in specific parts and get a short onset of action of the drugs.

According to another aspect of the present disclosure, crystalline form CS7 of apabetalone is provided (hereinafter referred to as Form CS7). Said Form CS7 is an anhydrate.

The X-ray powder diffraction pattern of Form CS7 shows characteristic peaks at 2theta values of $5.9^{\circ} \pm 0.2^{\circ}$, $6.7^{\circ} \pm 0.2^{\circ}$, $10.7^{\circ} \pm 0.2^{\circ}$ and $12.5^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS7 shows one or two or three characteristic peaks at 2theta values of $8.4^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2$ and $13.3^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS7 shows characteristic peaks at 2theta values of $8.4^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2$ and $13.3^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS7 shows one or more characteristic peaks at 2theta values of $16.0^{\circ} \pm 0.2^{\circ}$, $25.1^{\circ} \pm 0.2^{\circ}$, $15.0^{\circ} \pm 0.2^{\circ}$ and $21.8^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS7 shows characteristic peaks at 2theta values of $16.0^{\circ} \pm 0.2^{\circ}$, $25.1^{\circ} \pm 0.2^{\circ}$, $15.0^{\circ} \pm 0.2^{\circ}$ and $21.8^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS7 shows characteristic peaks at 2theta values of $5.9^{\circ} \pm 0.2^{\circ}$, $6.7^{\circ} \pm 0.2^{\circ}$, $10.7^{\circ} \pm 0.2^{\circ}$, $12.5^{\circ} \pm 0.2^{\circ}$, $8.4^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2$, $13.3^{\circ} \pm 0.2^{\circ}$, $16.0^{\circ} \pm 0.2^{\circ}$, $25.1^{\circ} \pm 0.2^{\circ}$, $15.0^{\circ} \pm 0.2^{\circ}$, $21.8^{\circ} \pm 0.2^{\circ}$ and $24.5^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction pattern of Form CS7 is substantially as depicted in FIG 6A.

According to another aspect of the present disclosure, a process for preparing Form CS7 of apabetalone is also provided. The process comprises:

Dissolving step: Dissolving apabetalone solid in a solvent of halohydrocarbons to obtain a clear solution.

Precipitation step: The solid is induced by liquid vapor diffusion. The obtained solid is Form CS7 of apabetalone.

5 Said precipitation step comprises: Storing the prepared solution in an open glass vial. And then putting the vial into another vial containing a solvent of ketones to obtain a solid. The obtained solid is Form CS7 of apabetalone.

Preferably, said halohydrocarbon includes chloroform, and said ketone includes methyl isobutyl ketone.

0 Form CS7 of the present disclosure has the following advantages:

1) At present, there is no patent or literature about apabetalone's crystalline forms. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS7, which is suitable for drug development.

2) Form CS7 has good stability. Form CS7 is stable for at least 4 weeks when stored under the
5 conditions of 25 °C/60% RH and 40 °C/75% RH. The stability of crystalline form is very important for drug development. Form CS7 has excellent stability, which can ensure that the quality of the drug will not be affected as polymorphic transition will not occur in the process of preparation, transportation and storage. It is of great significance to ensure the efficacy and safety of the drug and prevent the occurrence of adverse drug reactions.

20 According to another aspect of the present disclosure, crystalline form CS9 of apabetalone is provided (hereinafter referred to as Form CS9). Said Form CS9 is an anhydrate.

The X-ray powder diffraction pattern of Form CS9 shows characteristic peaks at 2theta values of $7.3^{\circ} \pm 0.2^{\circ}$, $9.9^{\circ} \pm 0.2^{\circ}$ and $17.0^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS9 shows one or two or three
25 characteristic peaks at 2theta values of $13.4^{\circ} \pm 0.2^{\circ}$, $3.9^{\circ} \pm 0.2^{\circ}$ and $12.8^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS9 shows characteristic peaks at 2theta values of $13.4^{\circ} \pm 0.2^{\circ}$, $3.9^{\circ} \pm 0.2^{\circ}$ and $12.8^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS9 shows one or more characteristic peaks at 2theta values of $12.1^{\circ} \pm 0.2^{\circ}$, $24.9^{\circ} \pm 0.2^{\circ}$, $22.5^{\circ} \pm 0.2^{\circ}$ and $24.2^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray
30 powder diffraction pattern of Form CS9 shows characteristic peaks at 2theta values of $12.1^{\circ} \pm 0.2^{\circ}$, $24.9^{\circ} \pm 0.2^{\circ}$, $22.5^{\circ} \pm 0.2^{\circ}$ and $24.2^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS9 shows characteristic peaks at 2theta values of $7.3^{\circ}\pm 0.2^{\circ}$, $9.9^{\circ}\pm 0.2^{\circ}$, $17.0^{\circ}\pm 0.2^{\circ}$, $13.4^{\circ}\pm 0.2^{\circ}$, $3.9^{\circ}\pm 0.2^{\circ}$, $12.8^{\circ}\pm 0.2^{\circ}$, $12.1^{\circ}\pm 0.2^{\circ}$, $24.9^{\circ}\pm 0.2^{\circ}$, $22.5^{\circ}\pm 0.2^{\circ}$, $24.2^{\circ}\pm 0.2^{\circ}$ and $6.0^{\circ}\pm 0.2^{\circ}$ using CuK α radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction pattern of Form CS9 is substantially as depicted in FIG 7A.

According to another aspect of the present disclosure, a process for preparing Form CS9 of apabetalone is also provided. The process comprises:

Dissolving step: Dissolving apabetalone solid in a mixture of ethers and alcohols or a mixture of halohydrocarbons and alcohols to obtain a clear solution.

Precipitation step: Evaporating the prepared solution at room temperature to obtain a solid. The obtained solid is Form CS9 of apabetalone.

Furthermore, said ether includes tetrahydrofuran, said alcohol includes isopropanol; said halohydrocarbons include dichloromethane and chloroform. Said volume ratio of ether and alcohol is 1:1; said volume ratio of halohydrocarbon and alcohol is 4:1.

Form CS9 of the present disclosure has the following advantages:

1) At present, there is no patent or literature about apabetalone's crystalline forms. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS9, which is suitable for drug development.

2) Form CS9 has good stability. Form CS9 is stable for at least 10 months when stored under the condition of 25 °C/60% RH and 40 °C/75% RH. The stability of crystalline form is very important for drug development. Form CS9 has excellent stability, which can ensure that the quality of the drug will not be affected as polymorphic transition will not occur in the process of preparation, transportation and storage. It is of great significance to ensure the efficacy and safety of the drug and prevent the occurrence of adverse drug reactions.

3) Form CS9 of the present disclosure is almost non hygroscopic. The weight gain of Form CS9 of the present disclosure at 80% RH is 0.18% and the crystalline form doesn't change after DVS test. For almost non hygroscopic crystalline forms, it is not necessary to control the environmental humidity in production process. There is no special strict requirements for packaging and storage conditions. It saves costs and is easy to industrialize production and long-term storage of drugs. Because the storage conditions are not demanding, the cost of material storage and quality control will be greatly reduced, which has strong economic value and is more suitable for medicinal use.

According to another aspect of the present disclosure, crystalline form CS11 of apabetalone is provided (hereinafter referred to as Form CS11). Said Form CS11 is a hydrate.

The X-ray powder diffraction pattern of Form CS11 shows characteristic peaks at 2theta values of $7.8^{\circ} \pm 0.2^{\circ}$, $8.8^{\circ} \pm 0.2^{\circ}$, $9.7^{\circ} \pm 0.2^{\circ}$ and $13.6^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS11 shows one or two or three characteristic peaks at 2theta values of $4.4^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$ and $21.6^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS11 shows characteristic peaks at 2theta values of $4.4^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$ and $21.6^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS11 shows one or two characteristic peaks at 2theta values of $13.0^{\circ} \pm 0.2^{\circ}$ and $15.3^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS11 shows characteristic peaks at 2theta values of $13.0^{\circ} \pm 0.2^{\circ}$ and $15.3^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS11 shows characteristic peaks at 2theta values of $7.8^{\circ} \pm 0.2^{\circ}$, $8.8^{\circ} \pm 0.2^{\circ}$, $9.7^{\circ} \pm 0.2^{\circ}$, $13.6^{\circ} \pm 0.2^{\circ}$, $4.4^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$, $21.6^{\circ} \pm 0.2^{\circ}$, $13.0^{\circ} \pm 0.2^{\circ}$, $15.3^{\circ} \pm 0.2^{\circ}$, $22.7^{\circ} \pm 0.2^{\circ}$, $7.6^{\circ} \pm 0.2^{\circ}$ and $17.6^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction pattern of Form CS11 is substantially as depicted in FIG 8A.

According to another aspect of the present disclosure, a process for preparing Form CS11 of apabetalone is also provided. The process comprises:

Method 1: Dissolving step: Dissolving apabetalone solid in a solvent of alcohols to obtain a clear solution.

Precipitation step: Adding the prepared solution to water or adding water to the prepared solution to obtain a solid. The obtained solid is Form CS11 of apabetalone.

Method 2: Dissolving step: Dissolving apabetalone solid in a solvent of halohydrocarbons or a mixture of alcohols and ketones or a mixture of alcohols and arenes to obtain a clear solution.

Precipitation step: Evaporating the prepared solution at room temperature to obtain a solid. The obtained solid is Form CS11 of apabetalone.

Furthermore, said alcohol of method 1 includes methanol, said alcohol of method 2 includes methanol; said ketones include acetone and methyl isobutyl ketone. Said volume ratio of alcohol and ketone is 1:1 to 2:1; said volume ratio of alcohol and arene is 4:1.

Form CS11 of the present disclosure has following advantages:

1) At present, there is no patent or literature about apabetalone's crystalline forms. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS11, which is suitable for drug development.

2) Form CS11 has good stability. Form CS11 is stable for at least 6 weeks when stored under the condition of 25 °C/60% RH and 40 °C/75% RH. Form CS11 has excellent stability, which can ensure that the quality of the drug will not be affected as polymorphic transition will not occur in the process of preparation, transportation and storage. It is of great significance to ensure the efficacy and safety of the drug and prevent the occurrence of adverse drug reactions.

3) Form CS11 has excellent solubility, especially, the solubility in SGF of Form CS11 is as high as 0.71 mg/mL at 1 hour, and real-time solubilities in FeSSIF at 1 hour, 4 hours and 24 hours are all higher than 0.26 mg/mL. Highly soluble crystalline forms are conducive to improving drug's blood concentration and bioavailability.

According to another aspect of the present disclosure, crystalline form CS4 of apabetalone is provided (hereinafter referred to as Form CS4). Said Form CS4 is an anhydrate.

The X-ray powder diffraction pattern of Form CS4 shows characteristic peaks at 2theta values of $9.1^{\circ} \pm 0.2^{\circ}$, $14.5^{\circ} \pm 0.2^{\circ}$, $23.5^{\circ} \pm 0.2^{\circ}$ and $24.2^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Furthermore, the X-ray powder diffraction pattern of Form CS4 shows one or two or three characteristic peaks at 2theta values of $10.3^{\circ} \pm 0.2^{\circ}$, $25.0^{\circ} \pm 0.2^{\circ}$ and $26.3^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS4 shows characteristic peaks at 2theta values of $10.3^{\circ} \pm 0.2^{\circ}$, $25.0^{\circ} \pm 0.2^{\circ}$ and $26.3^{\circ} \pm 0.2^{\circ}$.

Furthermore, the X-ray powder diffraction pattern of Form CS4 shows one or two or three characteristic peaks at 2theta values of $10.8^{\circ} \pm 0.2^{\circ}$, $11.6^{\circ} \pm 0.2^{\circ}$ and $19.5^{\circ} \pm 0.2^{\circ}$. Preferably, the X-ray powder diffraction pattern of Form CS4 shows characteristic peaks at 2theta values of $10.8^{\circ} \pm 0.2^{\circ}$, $11.6^{\circ} \pm 0.2^{\circ}$ and $19.5^{\circ} \pm 0.2^{\circ}$.

In a preferred embodiment, the X-ray powder diffraction pattern of Form CS4 shows characteristic peaks at 2theta values of $9.1^{\circ} \pm 0.2^{\circ}$, $14.5^{\circ} \pm 0.2^{\circ}$, $23.5^{\circ} \pm 0.2^{\circ}$, $24.2^{\circ} \pm 0.2^{\circ}$, $10.3^{\circ} \pm 0.2^{\circ}$, $25.0^{\circ} \pm 0.2^{\circ}$, $26.3^{\circ} \pm 0.2^{\circ}$, $10.8^{\circ} \pm 0.2^{\circ}$, $11.6^{\circ} \pm 0.2^{\circ}$ and $19.5^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

Without any limitation being implied, in a specific embodiment the X-ray powder diffraction pattern of Form CS4 is substantially as depicted in FIG 9A.

According to another aspect of the present disclosure, a process for preparing Form CS4 of apabetalone is also provided. The process comprises: Heating Form CS11 of apabetalone to 200-220 °C, and then Form CS4 of apabetalone was obtained.

Form CS4 of the present disclosure has the following advantages:

5 1) At present, there is no patent or literature about apabetalone's crystalline forms. The inventor of the present disclosure has addressed this issue after a lot of experimental research and found Form CS4, which is suitable for drug development.

2) Form CS4 has good long-term stability and mechanical stability. Form CS4 is stable for at least 10 months when stored under the condition of 25 °C/60%RH and 40 °C/75%RH. The stability of
0 crystalline form is very important for drug development. Form CS4 has excellent stability, which can ensure that the the quality of the drug will not be affected as polymorphic transition will not

occur in the process of preparation, transportation and storage. It is of great significance to ensure the efficacy and safety of the drug and prevent the occurrence of adverse drug reactions. In addition, Form CS4 doesn't change after manual grinding, which indicates that Form CS4 has good mechanical stability. The grinding of APIs is usually needed in production process. Good grinding stability can reduce the risk of crystallinity change and polymorphic transition of APIs during production process.

3) Form CS4 of the present disclosure is almost non hygroscopic. The weight gain of Form CS4 of the present disclosure at 80% RH is 0.12% and the crystalline form doesn't change after DVS test. For almost non hygroscopic crystalline forms, it is not necessary to control the environmental humidity in production process. There is no special strict requirements for packaging and storage conditions. It saves costs and is easy to industrialize production and long-term storage of drugs. Because the storage conditions are not demanding, the cost of material storage and quality control will be greatly reduced, which has strong economic value and is more suitable for medicinal use.

In the processes for preparation of Form CS2, Form CS8, Form CS13, Form CS20, Form CS1, Form CS7, Form CS9, Form CS11 and Form CS4 of the present disclosure:

Said "room temperature" is not an specific temperature value and refers to 10-30 °C

Said "evaporating" is accomplished by using a conventional method in the field. Slow evaporation is accomplished in a container covered by sealing film with pinholes. Rapid evaporation is accomplished in an open container.

Said "stirring" is accomplished by using a conventional method in the field such as magnetic stirring or mechanical stirring. The stirring speed is 50 to 1800 r/min, preferably is 300 to 900 r/min.

Said "cooling " is accomplished by using conventional methods in the field such as slow cooling and rapid cooling. Slow cooling is usually accomplished at a speed of 0.1 °C/min. Rapid cooling is usually accomplished by transferring the sample directly from high temperature environment to low temperature environment.

In the present disclosure, "crystal" or "crystalline form" refers to the crystal or the crystalline form being identified by the X-ray diffraction pattern shown herein. Those skilled in the art are able to understand that physicochemical properties discussed herein can be characterized. The experimental errors depend on the instrument conditions, the sampling processes and the purity of samples. In particular, those skilled in the art generally know that the X-ray diffraction pattern

typically varies with the experimental conditions. It is necessary to point out that, the relative intensity of the diffraction peaks in the X-ray diffraction pattern may also vary with the experimental conditions; therefore, the order of the diffraction peak intensities cannot be regarded as the sole or decisive factor. In fact, the relative intensity of the diffraction peaks in the X-ray powder diffraction pattern is related to the preferred orientation of the crystals, and the diffraction peak intensities shown herein are illustrative and identical diffraction peak intensities are not required. In addition, the experimental error of the diffraction peak position is usually 5% or less, and the error of these positions should also be taken into account. An error of $\pm 0.2^\circ$ is usually allowed. In addition, due to experimental factors such as sample thickness, the overall offset of the diffraction peak is caused, and a certain offset is usually allowed. Thus, it will be understood by those skilled in the art that a crystalline form of the present disclosure is not necessarily to have the exactly same X-ray diffraction pattern of the example shown herein. As used herein, "the same XRPD pattern" does not mean absolutely the same, the same peak positions may differ by $\pm 0.2^\circ$ and the peak intensity allows for some variability. Any crystalline forms whose X-ray diffraction patterns have the same or similar characteristic peaks should be within the scope of the present disclosure. Those skilled in the art can compare the patterns shown in the present disclosure with that of an unknown crystalline form in order to identify whether these two groups of patterns reflect the same or different crystalline forms.

In some embodiments, Form CS2, Form CS8, Form CS13, Form CS20, Form CS1, Form CS7, Form CS9, Form CS11 and Form CS4 of the present disclosure are pure and substantially free of any other crystalline forms. In the present disclosure, the term "substantially free" when used to describe a novel crystalline form, it means that the content of other crystalline forms in the novel crystalline form is less than 20% (w/w), specifically less than 10% (w/w), more specifically less than 5% (w/w) and further more specifically less than 1% (w/w).

It should be noted that the number and the number range should not be understood as the number or number range themselves only. It should be understood by those skilled in the art that the specific number can be shifted at specific technical environment without departing from the spirit and principle of the present disclosure. In the present disclosure, the number of shift ranges expected by one of skilled in the art is represented by the term "about".

The present disclosure also provides a mixed crystalline form of apabetalone, which contains more than two crystalline forms of Form CS2, Form CS8, Form CS13, Form CS20, Form CS1,

Form CS7, Form CS9, Form CS11 and Form CS4 in any proportion.

In addition, the present disclosure provides a pharmaceutical composition, said pharmaceutical composition comprises a therapeutically and/or prophylactically effective amount of Form CS2, Form CS8, Form CS13, Form CS20, Form CS1, Form CS7, Form CS9, Form CS11 and Form CS4 and at least one pharmaceutically acceptable carriers, diluents or excipients.

Furthermore, Form CS2, Form CS8, Form CS13, Form CS20, Form CS1, Form CS7, Form CS9, Form CS11 and Form CS4 can be used for preparing drugs treating cardiovascular, cholesterol or lipid-related disorders.

Furthermore, Form CS2, Form CS8, Form CS13, Form CS20, Form CS1, Form CS7, Form CS9, Form CS11 and Form CS4 can be used for preparing drugs preventing and treating cardiovascular diseases and diabetes, especially for preparing drugs treating atherosclerosis, acute coronary syndrome and predecessor diabetes.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1A shows an XRPD pattern of Form CS2 according to example 1 of the present disclosure.

FIG. 1B shows a DSC curve of Form CS2 according to example 1 of the present disclosure.

FIG. 1C shows a TGA curve of Form CS2 according to example 1 of the present disclosure.

FIG. 2A shows an XRPD pattern of Form CS8 according to example 2 of the present disclosure.

FIG. 2B shows a DSC curve of Form CS8 according to example 2 of the present disclosure.

FIG. 2C shows a TGA curve of Form CS8 according to example 2 of the present disclosure.

FIG. 2D shows an XRPD pattern overlay of Form CS8 before and after being stored under 25 °C/60%RH for two weeks (from top to bottom: XRPD pattern before storage, XRPD pattern after being stored under 25 °C/60%RH for two weeks).

FIG. 2E shows a DVS plot of Form CS8 according to example 2 of the present disclosure.

FIG. 2F shows an XRPD pattern overlay of Form CS8 according to example 2 of the present disclosure before and after DVS test.

FIG. 3A shows an XRPD pattern of Form CS13 according to example 3 of the present disclosure.

FIG. 3B shows a DSC curve of Form CS13 according to example 3 of the present disclosure.

FIG. 3C shows a TGA curve of Form CS13 according to example 3 of the present disclosure.

FIG. 4A shows an XRPD pattern of Form CS20 according to example 4 of the present disclosure.

FIG. 4B shows a DSC curve of Form CS20 according to example 4 of the present disclosure.

FIG. 4C shows a TGA curve of Form CS20 according to example 4 of the present disclosure.

FIG. 4D shows a ^1H NMR spectrum of Form CS20 according to example 4 of the present disclosure.

FIG. 4E shows an XRPD pattern overlay of Form CS20 before and after being stored under 25 °C/60%RH and 40 °C/75%RH for two weeks (from top to bottom: XRPD pattern before storage, XRPD pattern after being stored under 25 °C/60%RH for two weeks, XRPD pattern after being stored under 40 °C/75%RH for two weeks).

FIG. 5A shows an XRPD pattern of Form CS1 according to example 5 method 1 of the present disclosure.

FIG. 5B shows a DSC curve of Form CS1 according to example 5 method 1 of the present disclosure.

FIG. 5C shows a TGA curve of Form CS1 according to example 5 method 1 of the present disclosure.

FIG. 5D shows an XRPD pattern of Form CS1 according to example 5 method 2 of the present disclosure.

FIG. 5E shows an XRPD pattern overlay of Form CS1 before and after being stored under 25 °C/60%RH and 40 °C/75%RH for 10 months (from top to bottom: XRPD pattern before storage, XRPD pattern after being stored under 25 °C/60%RH for 10 months, XRPD pattern after being stored under 40 °C/75%RH for 10 months).

FIG. 5F shows a DVS plot of Form CS1 according to example 5 of the present disclosure.

FIG. 5G shows an XRPD pattern overlay of Form CS1 according to example 5 of the present disclosure before and after DVS test.

FIG. 5H shows an XRPD pattern overlay of Form CS1 according to example 5 of the present disclosure before and after grinding.

FIG. 5I shows a dissolution rate profile of Form CS1 according to example 5 of the present disclosure in formulations.

FIG. 6A shows an XRPD pattern of Form CS7 according to example 6 of the present disclosure.

FIG. 6B shows a DSC curve of Form CS7 according to example 6 of the present disclosure.

FIG. 6C shows a TGA curve of Form CS7 according to example 6 of the present disclosure

FIG. 6D shows an XRPD pattern overlay of Form CS7 before and after being stored under 25 °C/60%RH and 40 °C/75%RH for 4 weeks and 80 °C for one day (from top to bottom: XRPD pattern before storage, XRPD pattern after being stored under 25 °C/60%RH for 4 weeks, XRPD pattern after being stored under 40 °C/75%RH for 4 weeks, XRPD pattern after being stored under 80 °C for one day).

FIG. 6E shows an XRPD pattern overlay of Form CS7 according to example 6 of the present disclosure before and after DVS test.

FIG. 7A shows an XRPD pattern of Form CS9 according to example 7 of the present disclosure.

FIG. 7B shows a DSC curve of Form CS9 according to example 7 of the present disclosure.

FIG. 7C shows a TGA curve of Form CS9 according to example 7 of the present disclosure.

FIG. 7D shows an XRPD pattern overlay of Form CS9 before and after being stored under 25 °C/60%RH and 40 °C/75%RH for 10 months (from top to bottom: XRPD pattern before storage, XRPD pattern after being stored under 25 °C/60%RH for 10 months, XRPD pattern after being stored under 40 °C/75%RH for 10 months).

FIG. 7E shows a DVS plot of Form CS9 according to example 7 of the present disclosure.

FIG. 7F shows an XRPD pattern overlay of Form CS9 according to example 7 of the present disclosure before and after DVS test

FIG. 8A shows an XRPD pattern of Form CS11 according to example 8 method 1 of the present disclosure

FIG. 8B shows a DSC curve of Form CS11 according to example 8 method 1 of the present disclosure

FIG. 8C shows a TGA curve of Form CS11 according to example 8 method 1 of the present disclosure

FIG. 8D shows an XRPD pattern of Form CS11 according to example 8 method 2 of the present disclosure

FIG. 8E shows an XRPD pattern overlay of Form CS11 before and after being stored under 25 °C/60%RH and 40 °C/75%RH for 6 weeks (from top to bottom: XRPD pattern before storage, XRPD pattern after being stored under 25 °C/60%RH for 6 weeks, XRPD pattern after being stored under 40 °C/75%RH for 6 weeks).

FIG. 9A shows an XRPD pattern of Form CS4 according to example 9 of the present disclosure

FIG. 9B shows a DSC curve of Form CS4 according to example 9 of the present disclosure

FIG. 9C shows a TGA curve of Form CS4 according to example 9 of the present disclosure

FIG. 9D shows an XRPD pattern overlay of Form CS4 before and after being stored under 25

°C/60%RH and 40 °C/75%RH for 10 months (from top to bottom: XRPD pattern before storage,

5 XRPD pattern after being stored under 25 °C/60%RH for 10 months, XRPD pattern after being stored under 40 °C/75%RH for 10 months).

FIG. 9E shows a DVS plot of Form CS4 according to example 9 of the present disclosure

FIG. 9F shows an XRPD pattern overlay of Form CS4 according to example 9 of the present disclosure before and after DVS test.

10 FIG. 9G shows an XRPD pattern overlay of Form CS4 according to example 9 of the present disclosure before and after grinding.

DETAILED DESCRIPTION

The present disclosure is further illustrated by the following examples which describe the preparation and use of the crystalline forms of the present disclosure in detail. It is obvious to
15 those skilled in the art that many changes in the materials and methods can be accomplished without departing from the scope of the present disclosure.

The abbreviations used in the present disclosure are explained as follows:

XRPD: X-ray Powder Diffraction

HPLC: High Performance Liquid Chromatography

20 DVS: Dynamic Vapor Sorption

TGA: Thermal Gravimetric Analysis

¹H NMR: Proton Nuclear Magnetic Resonance

RH: Relative Humidity

Instruments and methods used for data collection:

25 X-ray powder diffraction patterns in the present disclosure were acquired by a Panalytical Empyrean X-ray powder diffractometer. The parameters of the X-ray powder diffraction method of the present disclosure were as follows:

X-ray Reflection: Cu, Kα

Kα1 (Å): 1.5400598; Kα2 (Å): 1.544426

30 Kα2/Kα1 intensity ratio: 0.50

Voltage: 45 (kV)

Current: 40 (mA)

Scan range: from 3.0 degree to 40.0 degree

Differential scanning calorimetry (DSC) data in the present disclosure were acquired by a TA Instruments Q200 MDSC. Instrument control software is Thermal Advantage, and analysis software is Universal Analysis.

Heating rate: 10 °C/min

Purge gas: nitrogen

Thermal gravimetric analysis (TGA) data in the present disclosure were acquired by a TA Instruments Q500 TGA. Instrument control software is Thermal Advantage, and analysis software is Universal Analysis.

Heating rate: 10 °C/ min

Purge gas: nitrogen

Proton nuclear magnetic resonance spectrum data (¹H NMR) were collected from a Bruker Avance II DMX 400M HZ NMR spectrometer. 1-5 mg of sample was weighed, and dissolved in 0.5 mL of deuterated dimethyl sulfoxide to obtain a solution with a concentration of 2-10 mg/mL.

Dynamic Vapor Sorption (DVS) was measured via a SMS (Surface Measurement Systems Ltd.) intrinsic DVS instrument. Its control software is DVS- Intrinsic control software, and its analysis software is DVS-Intrinsic Analysis software. Typical Parameters for DVS test are as follows:

Temperature: 25 °C

Gas and flow rate: N₂, 200 mL/min

dm/dt: 0.002%/min

RH range: 0% RH to 95% RH

High Performance Liquid Chromatography (HPLC) data in the present disclosure were collected from an Agilent 1260 with Diode Array Detector (DAD). The HPLC method parameters for purity test in the present disclosure are as follows:

1. Column: Waters XBridge C18 150×4.6mm, 5 μm

2. Mobile Phase: A: 0.1% TFA in H₂O

B: 0.1% TFA in Acetonitrile

Gradient:

Time (min)	%B
0.0	20
5.0	30
6.0	80
8.0	80
8.1	20
10.0	20

3. Flow rate: 1.6 mL/min
4. Injection Volume: 5 μ L
5. Detection wavelength: 260 nm
6. Column Temperature: 40 $^{\circ}$ C
7. Diluent: Acetonitrile/H₂O (v/v, 1/1)

The dissolution testing in the present disclosure was performed on an Agilent 708-DS.

Unless otherwise specified, the following examples were conducted at room temperature.

Raw materials of apabetalone used in the following examples were prepared by methods disclosed in CN101641339B.

Example 1

Preparation of Form CS2

Certain amount of apabetalone was weighed and dissolved in corresponding solvent of Table 1.1 at T₁. The solution was filtered and cooled to T₂ slowly or rapidly. When precipitation occurred, solids was obtained after centrifugation and drying.

Table 1.1

Sample No.	Weight (mg)	Solvent (v/v)	Volume (mL)	Cooling rate	T ₁ ($^{\circ}$ C)	T ₂ ($^{\circ}$ C)
1-a	19.7	Methanol	1.0	Rapid	50	-20
1-b	10.2	Methanol/methyl isobutyl ketone (1:1)	1.0	Slow	50	5
1-c	10.2	Methanol/methyl isobutyl ketone (1:1)	1.0	Rapid	50	-20
1-d	20.2	Methanol/toluene (1:1)	1.0	Slow	50	5
1-e	20.2	Methanol/toluene (1:1)	1.0	Rapid	50	-20
1-f	11.3	Methanol/2-methyl tetrahydrofuran (1:1)	1.0	Slow	50	5
1-g	11.3	Methanol/2-methyl tetrahydrofuran (1:1)	1.0	Rapid	50	-20
1-h	10.7	Tetrahydrofuran/methyl tert-butyl ether (2:1)	3.0	Slow	50	5

1-i	19.9	Methanol	1.0	Rapid	100	-20
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Sample 1-a to 1-i were confirmed to be Form CS2 by XRPD. Sample 1-i was selected for characterization. The XRPD pattern is substantially as depicted in Figure 1A, and the XRPD data are listed in Table 1.2. The DSC curve of Form CS2 is substantially as depicted in Figure 1B, which shows the first endothermic peak at around 60 °C, the second endothermic peak at around 113 °C, the first exothermic peak at around 126 °C and the third endothermic peak at around 232 °C. The TGA curve of Form CS2 is substantially as depicted in Figure 1C, which shows about 3.4% weight loss when heated to 100 °C.

Table 1.2

2 Theta	d spacing	Intensity %
5.08	17.39	30.31
6.57	13.46	100.00
8.85	10.00	34.56
11.54	7.67	41.68
13.31	6.65	8.72
15.28	5.80	14.13
20.22	4.39	7.21
23.10	3.85	4.08
25.35	3.51	2.44

Solubility study of Form CS2

The prepared apabetalone Form CS2 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 1.3.

Table 1.3

Solubility (mg/mL) Time (h)	SGF	FeSSIF
1	0.49	0.15
24	0.61	0.24

The above results show that Form CS2 of apabetalone has good solubility in SGF and FeSSIF, especially in SGF, the solubility at 24 h is as high as 0.61 mg/mL. Polymorph with high solubility is conducive to increasing the blood concentration of drugs in human body and improves the

bioavailability of drugs, which is of great significance for drug research.

Example 2

Preparation of Form CS8

Certain amount of apabetalone was weighed and dissolved in corresponding solvent of Table 2.1 at T₁. The solution was filtered and cooled to T₂ slowly or rapidly. When precipitation occurred, solids was obtained after centrifugation and drying.

Table 2.1

Sample No.	Weight (mg)	Solvent (v/v)	Volume (mL)	Cooling rate	T ₁ (°C)	T ₂ (°C)
2-a	9.8	Dichloromethane	1.5	Slow	50	5
2-b	9.8	Dichloromethane	1.5	Rapid	50	-20
2-c	20.8	Dichloromethane/isoprop anol (4:1)	1.0	Slow	50	5
2-d	20.8	Dichloromethane/isoprop anol (4:1)	1.0	Rapid	50	-20

Sample 2-a to 2-d were confirmed to be Form CS8 by XRPD. Sample 2-a was selected for characterization. The XRPD pattern is substantially as depicted in Figure 2A, and the XRPD data are listed in Table 2.2. The DSC curve of Form CS8 is substantially as depicted in Figure 2B, which shows the first exothermic peak at around 177 °C, the first endothermic peak at around 211 °C, the second exothermic peak at around 217 °C and the second endothermic peak at around 230 °C. The TGA curve of Form CS8 is substantially as depicted in Figure 2C, which shows about 0.4% weight loss when heated to 120 °C.

Table 2.2

2 Theta	d spacing	Intensity%
4.23	20.89	3.93
7.83	11.30	60.72
9.74	9.08	6.55
10.58	8.36	5.86
11.36	7.79	16.82
13.11	6.75	16.65
13.55	6.54	40.35
18.00	4.93	10.44
19.29	4.60	9.65
20.17	4.40	14.20

22.47	3.96	25.75
23.91	3.72	100.00
24.22	3.67	32.56
24.87	3.58	11.65
25.88	3.44	14.42
26.18	3.40	17.03
28.16	3.17	14.65
31.72	2.82	2.11

Stability study of Form CS8

Stability Study of apabetalone Form CS8

Form CS8 was placed in a constant temperature and humidity chamber at 25 °C/60% RH for 2 weeks in open dish. Crystalline form of the sample were tested by XRPD and impurity of the sample were checked. The results are shown in Figure 2D (From top to bottom: XRPD pattern of Form CS8 before and after being stored under 25 °C/60% RH for 2 weeks.)

No obvious form change and purity decrease was observed for Form CS8 after being stored under 25 °C/60% RH for 2 weeks. The results show that Form CS8 has good stability.

Solubility study of Form CS8

The prepared Form CS8 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 2.3.

Table 2.3

Solubility (mg/mL) Time (h)	SGF	FeSSIF
	0.43	0.28

The above results show that Form CS8 of apabetalone has good solubility in SGF and FeSSIF. Polymorph with high solubility is beneficial to increase the blood concentration of drugs in human body and improve the bioavailability of drugs, which is of great significance for drug research.

Hygroscopicity of Form CS8

Dynamic vapor sorption (DVS) was applied to test hygroscopicity of Form CS8 in the present

disclosure with about 10 mg of sample. The result is listed in Table 2.4. The DVS plot of Form CS8 is substantially as depicted in Figure 2E.

Table 2.4

Form	Weight Gain under 80% Relative Humidity
Form CS8	0.34%

Description and definition of hygroscopicity (Chinese Pharmacopoeia 2015 edition appendix 9103 drug hygroscopic test guidelines, test at 25 °C+/-1 °C, 80% RH.).

—deliquescent: Sufficient water is absorbed to form a liquid;

—very hygroscopic: Increase in mass is equal to or greater than 15 percent;

—hygroscopic: Increase in mass is less than 15 percent and equal to or greater than 2 percent;

—slightly hygroscopic: Increase in mass is less than 2 percent and equal to or greater than 0.2 percent.

—non hygroscopic or almost non hygroscopic: Increase in mass is less than 0.2%.

The results indicates that the weight gain of Form CS8 under 80% RH is 0.34%. According to the definition of hygroscopicity, Form CS8 is slightly hygroscopic. The XRPD pattern of Form CS8 after DVS test is shown in Figure 2F. No form change was observed for Form CS8 after DVS test, which indicates that Form CS8 is stable under the influence of humidity.

Form CS8 of the present disclosure shows low hygroscopicity and can avoid issues such as crystal instability in the process of drug preparation and/or storage, as well as the unprocessability of the preparation caused by external factors such as environmental moisture, which is conducive to the accurate quantitative preparation and later transportation and storage.

Example 3

Preparation of Form CS13

Certain amount of apabetalone was weighed and dissolved in corresponding solvent shown in Table 3.1. The solution was filtered and evaporated slowly at room temperature with or without addition of polymer to obtain solid.

Table 3.1

Sample No.	Weight (mg)	Solvent (v/v)	Volume (mL)	Whether to add polymer (Y/N)	T(°C)
3-a	10.8	Tetrahydrofuran/water	1.3	N	25

		(4:1)			
3-b	10.5	Acetone/water (9:1)	3.3	N	25
3-c	10.8	Tetrahydrofuran/water (4:1)	1.3	Y	25
3-d	10.5	Acetone/water (9:1)	3.3	Y	25

Said polymer is a mixture of equal masses of polycaprolactone, polyethylene glycol, polymethyl methacrylate, sodium alginate and hydroxyethyl cellulose.

Sample 3-a to 3-d were confirmed to be Form CS13 by XRPD. Sample 3-d was selected for characterization. The XRPD pattern is substantially as depicted in Figure 3A, and the XRPD data are listed in Table 3.2. The DSC curve of Form CS13 is substantially as depicted in Figure 3B, which shows the first endothermic peak at around 70 °C, the second endothermic peak at around 86 °C, the third endothermic peak at around 170 °C, the fourth endothermic peak at around 205 °C, the first exothermic peak at around 207 °C and the fifth endothermic peak at around 230 °C. The TGA curve of Form CS13 is substantially as depicted in Figure 3C, which shows about 11.9% weight loss when heated to 100 °C.

Table 3.2

2 Theta	d spacing	Intensity%
5.09	17.35	48.06
6.38	13.85	22.94
7.76	11.39	4.83
8.54	10.35	12.25
10.17	8.70	0.71
11.48	7.71	3.89
12.54	7.06	100.00
13.29	6.66	3.59
15.30	5.79	1.21
15.56	5.69	2.14
16.01	5.54	4.65
17.13	5.18	47.37
19.15	4.64	0.27
20.41	4.35	1.35
20.84	4.26	4.59
22.11	4.02	0.59

23.43	3.80	3.37
25.67	3.47	8.18
26.75	3.33	1.24
27.06	3.29	3.65
28.72	3.11	1.18
29.56	3.02	1.42
30.89	2.90	0.43
32.26	2.77	1.36
34.67	2.59	0.58
37.28	2.41	2.20
38.97	2.31	0.89

Solubility of Form CS13

The prepared Form CS13 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, 4 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 3.3.

Table 3.3

Solubility (mg/mL) Time (h)	SGF	FeSSIF
1	0.35	0.13
4	0.31	0.14
24	0.33	0.14

The above results show that Form CS13 of apabetalone has good solubility in SGF and FeSSIF.

Example 4

Preparation of Form CS20

Certain amount of apabetalone was weighed and dissolved in corresponding solvent of Table 4.1. The solution was filtered and evaporated slowly at room temperature to obtain a solid.

Table 4.1

Sample No.	Weight (mg)	Solvent (v/v)	Volume (mL)	T(°C)
4-a	10.5	Acetonitrile/acetic acid (9:1)	2.0	25
4-b	10.3	Ethyl acetate/acetic acid (4:1)	1.7	25

Sample 4-a to 4-b were confirmed to be Form CS20 of apabetalone by XRPD. Sample 4-a was

selected for characterization. The XRPD pattern is substantially as depicted in Figure 4A, and the XRPD data are listed in Table 4.2. The DSC curve of Form CS20 is substantially as depicted in Figure 4B, which shows the first endothermic peak at around 123 °C, the second endothermic peak at around 185 °C, the third endothermic peak at around 201 °C, the fourth endothermic peak at around 207 °C and the fifth endothermic peak at around 230 °C. The TGA curve of Form CS13 is substantially as depicted in Figure 4C, which shows about 14.9% weight loss when heated to 160 °C. The ¹H NMR spectrum of Form CS20 is substantially as depicted in Figure 4D, and the corresponding data are: ¹H NMR (400 MHz, DMSO) δ 11.83 (s, 1H), 7.89 (s, 2H), 6.74 (d, J = 2.3 Hz, 1H), 6.52 (d, J = 2.3 Hz, 1H), 4.91 (s, 1H), 3.89 (s, 3H), 3.87 -3.81 (m, 5H), 3.72 (s, 2H), 2.31 (s, 6H), 1.90 (s, 3H).

Table 4.2

2 Theta	d spacing	Intensity%
5.59	15.81	21.68
7.20	12.28	4.52
8.39	10.53	100.00
9.37	9.44	22.64
11.26	7.86	25.84
11.67	7.58	8.95
13.52	6.55	83.84
13.88	6.38	7.82
14.48	6.12	11.79
14.92	5.94	3.42
16.92	5.24	11.24
18.88	4.70	30.97
19.57	4.54	1.85
20.15	4.41	12.22
20.56	4.32	12.06
22.30	3.99	5.29
22.76	3.91	10.73
24.10	3.69	2.41
24.38	3.65	4.37
25.26	3.53	1.52
25.93	3.44	3.39

26.33	3.38	13.70
26.78	3.33	3.08
28.11	3.17	2.02
28.81	3.10	3.75
34.30	2.61	1.52
35.68	2.52	2.00
36.85	2.44	2.14

Stability Study of Form CS20

Two samples of apabetalone Form CS20 were placed in constant temperature and humidity chambers at 25 °C/60% RH and 40 °C/75% RH for 2 weeks in open dish. Crystalline form of the sample were tested by XRPD and impurity of the sample were checked. The XRPD pattern overlay is substantially as depicted in Figure 4E (from top to bottom: XRPD pattern of Form CS20 before and after being stored under 25 °C/60% RH and 40 °C/75% RH for 2 weeks). No form change and obvious purity decrease was observed for Form CS20 after being stored at 25 °C/60% RH and 40 °C/75% RH for 2 weeks. It can be seen that Form CS20 has good stability.

Solubility study of Form CS20

The prepared apabetalone Form CS20 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, 4 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 4.3.

Table 4.3

Solubility (mg/mL) Time (h)	SGF	FeSSIF
1	0.44	0.19
4	0.41	0.16
24	0.51	0.22

The above results show that Form CS20 of apabetalone of the present disclosure has good solubility in SGF and FeSSIF. Polymorph with high solubility is beneficial to increase the blood concentration of drugs in human body and improve the bioavailability of drugs, which is of great significance for drug research.

Example 5

Preparation of Form CS1 (Method 1)

The process of preparing Form CS1 of apabetalone comprises the following steps:

Dissoving step: About 10 mg of apabetalone solid was dissolve in corresponding solvent of Table 5.1, and a clear solution was obtained by filtering.

Precipitation step: The prepared solution was added to corresponding anti-solvent of Table 5.1 or corresponding anti-solvent was added to the prepared solution with stirring until a lot of precipitation was observed. The solid was collected by centrifugation and drying. The obtained solid was Form CS1 of apabetalone.

Furthermore, the reaction conditions, composition and amount of solvents and anti-solvents of the preparing process for said Form CS1 of apabetalone are listed in Table 5.1.

Table 5.1

Sample No.	Weight (mg)	Solvent	Volume (mL)	Anti-solvent	Volume (mL)	Method
5-a	10.1	Tetrahydrofuran	2.125	n-Heptane	3.0	Anti-solvent addition
5-b	10.1	Tetrahydrofuran	2.125	n-Heptane	3.0	Reverse anti-solvent addition
5-c	10.1	Tetrahydrofuran	2.125	Methyl tertiary butyl ether	3.0	Anti-solvent addition
5-d	10.1	Tetrahydrofuran	2.125	Methyl tertiary butyl ether	3.0	Reverse anti-solvent addition
5-e	10.1	Tetrahydrofuran	2.125	Toluene	3.0	Reverse anti-solvent addition
5-f	10.1	Chloroform	1.375	Methyl tertiary butyl ether	2.0	Reverse anti-solvent addition
5-g	10.1	Chloroform	1.375	Toluene	3.0	Anti-solvent addition
5-h	9.9	Dimethyl sulfoxide	1.0	Water	3.0	Reverse anti-solvent addition
5-i	10.1	Dimethylacetamide	0.25	Acetonitrile	3.0	Reverse anti-solvent addition

Sample 5-a to 5-i were confirmed to be Form CS1 of apabetalone by XRPD. Sample 5-a was selected for characteriztion. The XRPD pattern is substantially as depicted in Figure 5A, and the XRPD data are listed in Table 5.2. The DSC curve of Form CS1 is substantially as depicted in Figure 5B, which shows the first endothermic peak at around 207 °C, the first exothermic peak at around 211 °C and the second endothermic peak at around 231 °C. The TGA curve of Form CS1 is substantially as depicted in Figure 5C, which shows about 0.7% weight loss when heated to 200

°C.

Table 5.2

2 Theta	d spacing	Intensity%
4.38	20.19	1.10
6.11	14.47	100.00
10.19	8.68	2.41
10.70	8.27	1.40
12.11	7.31	21.04
12.27	7.21	62.62
12.99	6.82	5.91
14.12	6.27	3.01
16.37	5.41	4.60
17.09	5.19	3.31
17.63	5.03	1.77
18.18	4.88	1.63
18.48	4.80	7.06
20.12	4.41	2.93
20.49	4.33	4.18
21.55	4.12	1.12
21.93	4.05	2.34
23.16	3.84	5.71
23.45	3.79	4.16
24.45	3.64	2.64
25.02	3.56	1.26
26.11	3.41	8.67
26.84	3.32	7.12
28.00	3.19	0.96
28.60	3.12	0.96
30.04	2.98	0.86
31.04	2.88	1.25
35.60	2.52	0.25
36.80	2.44	0.50
38.30	2.35	0.89

Preparation of Form CS1 (Method 2)

The process of preparing Form CS1 comprises the following steps:

Dissoving step: Apabetalone solid was dissolved in corresponding solvent of Table 5.3 at 50 °C to obtain a clear solution.

Precipitation step: The solution was cooled to -20-5 °C rapidly or slowly until solid precipatated.

The solid was collected by centrifugation and drying. The obtained solid was Form CS1 of apabetalone.

Said reaction conditions, solvent composition and solvent amount of the preparation method of apabetalone Form CS1 are shown in Table 5.3. Sample 5-j to 5-n were confirmed to be Form CS1 of apabetalone by XRPD. The XRPD pattern of sample 5-j is substantially as depicted in Figure 5D, and the XRPD data are listed in Table 5.4.

Table 5.3

Sample No.	Weight (mg)	Solvent (v/v)	Volume (mL)	Method	T(°C)
5-j	19.9	Tetrahydrofuran	1.0	Slow cooling	5
5-k	19.9	Tetrahydrofuran	1.0	Rapid cooling	-20
5-l	10.8	Acetone	3.0	Rapid cooling	-20
5-m	10.2	Ethyl acetate/acetone (1:1)	3.0	Rapid cooling	-20
5-n	10.7	Acetonitrile /N, N-dimethyl formamide (9:1)	3.0	Rapid cooling	-20

Table 5.4

2 Theta	d spacing	Intensity%
3.57	24.78	1.09
6.10	14.48	100.00
10.26	8.62	0.50
12.28	7.21	65.97
13.01	6.81	2.85
14.15	6.26	2.43
16.36	5.42	3.05
17.13	5.18	2.82
18.47	4.80	6.67
19.24	4.61	0.64
20.14	4.41	0.92

20.51	4.33	1.86
22.01	4.04	1.19
23.16	3.84	7.43
24.45	3.64	2.81
25.01	3.56	1.21
26.15	3.41	10.57
26.86	3.32	8.29
28.76	3.10	0.82
30.08	2.97	0.92
31.07	2.88	1.86
34.62	2.59	0.49
36.83	2.44	0.62
38.30	2.35	1.75

Long-term stability study of Form CS1

Two samples of apabetalone Form CS1 were placed in constant temperature and humidity chambers at 25 °C/60% RH and 40 °C/75% RH for 10 months in open dishes. The samples were characterized by XRPD and chemical impurity. The results are substantially as depicted in Figure 5E (from top to bottom: XRPD pattern of Form CS1 before and after being stored under 25 °C/60% RH and 40 °C/75% RH for 10 months) and Table 5.5.

No form change and obvious purity decrease was observed for Form CS1 after being stored at 25 °C/60% RH and 40 °C/75% RH for 10 months. It can be seen that Form CS1 has good stability and high purity.

Table 5.5

Condition	1 week	2 weeks	5 weeks	10 months
25 °C/60%RH	99.27	99.23	99.25	99.24
40 °C/75%RH	99.20	99.20	99.25	99.21

Solubility study of Form CS1

The prepared solid of apabetalone Form CS1 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, 4 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 5.6.

Table 5.6

Solubility (mg/mL) Time (h)	SGF	FeSSIF
1	0.31	0.14
4	0.29	0.11
24	0.33	0.14

The above results show that Form CS1 of apabetalone has good solubility in SGF and FeSSIF.

Hygroscopicity study of Form CS1

Dynamic vapor sorption (DVS) was applied to test hygroscopicity of Form CS1 of the present disclosure with about 10 mg of sample. The result is listed in Table 5.7. The DVS plot of Form CS1 is substantially as depicted in Figure 5F.

Table 5.7

Form	Weight Gain under 80% Relative Humidity
Form CS1	0.13%

The results showed that weight gain of Form CS1 under 80%RH is 0.13%. According to the hygroscopicity criteria, Form CS1 is almost non hygroscopic. The XRPD pattern of Form CS1 after DVS test is shown in Figure 5G. No form change was observed for Form CS1 before and after DVS test, which indicates that Form CS1 is stable under the influence of humidity.

Form CS1 of the present disclosure shows low hygroscopicity and can avoid issues such as crystal instability in the process of drug preparation and/or storage, as well as the unprocessability of the preparation caused by external factors such as environmental moisture, which is conducive to the accurate quantitative preparation and later transportation and storage.

Mechanical stability of Form CS1

Certain amount of Form CS1 was placed in a mortar and ground manually for 5 minutes. XRPD of the solid obtained was tested. The results are shown in Figure 5H.

The results showed that no form change and obvious crystallinity decrease was observed for Form CS1 of apabetalone under certain mechanical stress, and it can still maintain stable physical and chemical properties, which is suitable for drug preparation and storage. Grinding of API is usually needed in the process of formulation, and good grinding stability will reduce the risk of crystallinity decrease and transformation of solid form of API.

Formulation study of Form CS1

Certain amount of the API, microcrystalline cellulose, croscarmellose sodium according to Table

5.8 and 2 mg of magnesium stearate were weighted and blended for 2 minutes. Flakes were prepared using a manual tablet press at 5 KN pressure with a ϕ 20mm round tooling. The crushed flakes were manually sieved through 20 mesh sieve. 2 mg of magnesium stearate was added, and then the obtained powder was blended for 1 minute. The mixture was encapsulated into a 1# capsule shell, then the capsule was packed in 35cc HDPE (high density polyethylene) bottle (one capsule per bottle) containing 1g of desiccant. The bottle was then sealed by sealing machine. The composition of the formulation (per 200 mg) is shown in table 5.8 below. Form CS1 is stable in the preparations after testing.

Table 5.8

Component	Quantity (mg/capsule)
API	50
Microcrystalline cellulose	136
Croscarmellose sodium	10
Magnesium stearate	4

Dissolution test was performed on the obtained capsule. The conditions are as follows:

Medium: HCl (0.1 mol/L)

Method: Paddle

Volume: 900 mL

Speed: 75 rpm

Temperature: 37 °C

Dissolution results of Form CS1 are presented in Figure 5I, which indicates that Form CS1 possesses favorable dissolution. The released amount can reach more than 90% at 20 minutes. In drug product development, rapid dissolution rate can speed up the dissolution of the drug in body. By adjusting the excipients, it is possible to control the rapid action of drugs in specific parts and get a short onset of action of the drugs.

Example 6

Preparation of Form CS7

The process of preparing Form CS7 of apabetalone comprises the following steps:

Dissoving step: 5.3 mg of apabetalone solid was dissolved in 1 mL of chloroform, and a clear solution was obtained by filtering.

Precipitation step: The solution was added to a 3-mL glass vial. The vial was put into a 20-mL glass vial containing 5 mL of methyl isobutyl ketone for liquid vapor diffusion. Then the 20-mL vial was sealed and left at room temperature until solid precipitated. The solid was collected by

centrifugation and drying to obtain Form CS7 of apabetalone. The XRPD pattern of Form CS7 is substantially as depicted in Figure 6A, and the XRPD data are listed in Table 6.1.

The DSC curve shows the first endothermic peak at around 231 °C, which is substantially as depicted in Figure 6B. The TGA curve shows about 0.4% weight loss when heated to 160 °C, which is substantially as depicted in Figure 6C.

Table 6.1

2 Theta	d spacing	Intensity%
4.27	20.68	3.98
5.87	15.06	55.68
6.69	13.21	100
8.43	10.50	77.7
10.70	8.27	80.13
12.08	7.32	9.24
12.46	7.10	53.68
13.25	6.68	49.94
14.99	5.91	9.73
15.96	5.55	34.89
16.94	5.23	53.34
19.16	4.63	3.48
20.32	4.37	6.07
21.76	4.08	15.6
23.67	3.76	4.6
24.47	3.64	9.66
25.09	3.55	15.93
26.01	3.43	3.31
27.42	3.25	2.99
30.79	2.90	3.25
32.57	2.75	1.7
36.50	2.46	1.75
37.45	2.40	1.96

Stability study of Form CS7

Three samples of apabetalone Form CS7 were placed in constant temperature and humidity chambers at 25 °C/60% RH and 40 °C/75% RH for 4 weeks and 80 °C for 1 day in open dishes.

Crystalline form of the sample were tested by XRPD and impurity of the sample were checked.

The XRPD pattern overlay is substantially as depicted in Figure 6D (from top to bottom: XRPD pattern of Form CS7 before and after being stored under 25 °C/60% RH and 40 °C/75% RH for 4 weeks and 80 °C for 1 day).

No form change and obvious purity decrease was observed for Form CS7 after being stored at 25 °C/60% RH and 40 °C/75% RH for 4 weeks and 80 °C for 1 day. It can be seen that Form CS7 has good stability.

Solubility study of Form CS7

The prepared solid of apabetalone Form CS7 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, 4 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 6.2.

Table 6.2

Time (h) \ Solubility (mg/mL)	SGF	FeSSIF
1	0.38	0.19
4	0.42	0.17
24	0.58	0.29

The above results show that Form CS7 of apabetalone has good solubility in SGF and FeSSIF. Polymorph with high solubility is beneficial to increase the blood concentration of drugs in human body and improve the bioavailability of drugs, which is of great significance for drug research.

Hygroscopicity study of Form CS7

Dynamic vapor sorption (DVS) was applied to test hygroscopicity of Form CS7 of the present disclosure with about 10 mg of samples. The result is listed in Table 6.3.

Table 6.3

Form	Weight gain under 80% relative humidity
Form CS7	0.79%

The results showed that weight gain of Form CS7 under 80%RH is 0.79%. According to the hygroscopicity criteria, Form CS7 is slightly hygroscopic. The XRPD pattern of Form CS7 after DVS test was shown in Figure 6E. No form change was observed for Form CS7 before and after

DVS test, which indicates that Form CS7 has good humidity stability.

Form CS7 of the disclosure shows low hygroscopicity, which can well avoid issues such as crystal instability in the process of drug preparation and/or storage, as well as the unprocessability of the preparation caused by external factors such as environmental moisture, which is conducive to the accurate quantitative preparation and later transportation and storage.

Example 7

Preparation of Form CS9

The process of preparing Form CS9 comprises the following steps:

Dissolving step: About 10 mg of apabetalone solid was dissolved in corresponding solvent of Table 7.1, and a clear solution was obtained by filtering.

Precipitation step: The prepared solution was left at room temperature for slow evaporation until solid precipitated. The obtained solid was Form CS9 of Apabetalone.

Said reaction conditions, solvent composition and solvent amount of the preparation method of apabetalone Form CS9 are shown in Table 7.1. The samples 7-a to 7-c were confirmed to be Form CS9 by XRPD.

Table 7.1

Sample No.	Weight (mg)	Solvent (v/v)	Volume (mL)	Whether to add polymer or not (Y/N)	T (°C)
7-a	10.0	Tetrahydrofuran/isopropanol (1:1)	2.7	N	25
7-b	10.0	Dichloromethane/isopropanol (4:1)	1.7	N	25
7-c	10.7	Chloroform/acetonitrile (2:1)	1.7	Y	25

Sample 7-b was selected for characterization. The XRPD pattern is substantially as depicted in Figure 7A, and the XRPD data are listed in Table 7.2. The DSC curve is substantially as depicted in Figure 7B, which shows the first endothermic peak at around 203 °C. The TGA curve is substantially as depicted in Figure 7C, which shows about 0.9% weight loss when heated to 200 °C.

Table 7.2

2 Theta	d spcaing	Intensity%
3.92	22.55	17.55
5.99	14.75	7.28
7.25	12.19	64.22
7.94	11.13	6.24
9.92	8.91	24.30
12.09	7.32	8.02
12.78	6.93	17.69
13.36	6.63	22.02
13.53	6.54	10.74
14.48	6.12	2.15
16.98	5.22	19.92
20.01	4.44	7.41
20.88	4.26	5.94
21.42	4.15	21.02
21.77	4.08	16.45
22.00	4.04	14.36
22.53	3.95	100.00
23.02	3.86	25.09
23.95	3.72	34.15
24.19	3.68	58.53
24.86	3.58	16.90
25.83	3.45	25.83
26.11	3.41	24.07
26.53	3.36	7.10
27.03	3.30	6.89
27.81	3.21	7.53
28.51	3.13	10.88
30.61	2.92	3.05
31.17	2.87	2.73
32.29	2.77	2.32
32.80	2.73	2.65
33.67	2.66	2.03
36.63	2.45	1.42
37.12	2.42	1.68

Stablity study of Form CS9

Two samples of apabetalone Form CS9 were placed in constant temperature and humidity chambers at 25 °C/60% RH and 40 °C/75% RH for 10 months in open dishes. Crystalline form of the sample were tested by XRPD and impurity of the sample were tested to check the stability of Form CS9. The XRPD pattern overlay pattern is substantially as depicted in Figure 7D (from top to bottom: XRPD pattern of Form CS9 before and after being stored under 25 °C/60% RH and 40 °C/75% RH for 10 months).

No form change and obvious purity decrease was observed for Form CS9 after being stored at 25 °C/60% RH and 40 °C/75% RH for 10 months. It can be seen that Form CS9 has good stability.

Solubility study of Form CS9

The prepared solid of apabetalone Form CS9 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, 4 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 7.3.

Table 7.3

Solubility (mg/mL) Time (h)	SGF	FeSSIF
1	0.30	0.17
4	0.31	0.13
24	0.37	0.20

The above results show that Form CS9 of apabetalone has good solubility in SGF and FeSSIF.

Hygroscopicity study of Form CS9

Dynamic vapor sorption (DVS) was applied to test hygroscopicity of Form CS9 of the present disclosure with about 10 mg of sample. The result is listed in Table 7.4. The DVS plot of Form CS9 is substantially as depicted in Figure 7E. The XRPD pattern of Form CS9 after DVS test was shown in Figure 7F. No form change was observed for Form CS9 before and after DVS test, which indicates that Form CS9 is stable under the influence of humidity.

Table 7.4

Form	Weight gain under 80% relative humidity
Form CS9	0.18%

The results showed that weight gain of Form CS9 under 80%RH is 0.18%. According to the hygroscopicity criteria, Form CS9 is almost non hygroscopic. Form CS9 of the present disclosure shows low hygroscopicity, which can well avoid issues such as crystal instability in the process of drug preparation and/or storage, as well as the unprocessability of the preparation caused by external factors such as environmental moisture, which is conducive to the accurate quantitative preparation and later transportation and storage.

Example 8

Preparation of Form CS11 (Method 1)

The process of preparing Form CS11 comprises the following steps:

Dissolving step: about 10 mg of apabetalone solid was dissolved in 1.7 mL of methanol, and filtered to get a clear solution.

Precipitation step: The prepared solution was added to 3 mL of water or 3 mL of water was added to the prepared solution with stirring to obtain solid. Form CS11 was obtained by centrifugation and drying.

Said reaction conditions and amount of methanol and water of the preparation method of apabetalone Form CS11 are shown in Table 8.1. The samples 8-a to 8-b were confirmed to be Form CS11 by XRPD.

Table 8.1

Sample No.	Weight (mg)	Volume of methanol (mL)	Volume of water (mL)	Method
8-a	10.0	1.7	3	Antisolvent addition
8-b	10.0	1.7	3	Reverse anti-solvent addition

Sample 8-b was selected for characterization. The XRPD pattern is substantially as depicted in Figure 8A, and the XRPD data are listed in Table 8.2. The DSC curve is substantially as depicted in Figure 8B, which shows the first endothermic peak at around 49 °C, the second endothermic peak at around 206 °C, the first exothermic peak at 208 °C and the third endothermic peak at 230 °C. The TGA curve is substantially as depicted in Figure 8C, which shows about 3.0% weight loss when heated to 100 °C.

Table 8.2

2 Theta	d spacing	Intensity%
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4.36	20.25	22.75
7.61	11.61	35.17
7.80	11.33	71.58
8.79	10.06	75.43
9.70	9.12	100.00
10.11	8.75	11.17
13.02	6.80	8.06
13.55	6.54	23.03
15.32	5.79	8.84
15.70	5.64	7.37
16.87	5.26	11.53
17.64	5.03	10.29
18.67	4.75	1.01
19.55	4.54	3.26
20.40	4.35	5.32
20.99	4.23	5.64
21.58	4.12	21.02
22.72	3.91	9.87
23.67	3.76	3.85
25.65	3.47	5.77
26.85	3.32	2.83
27.34	3.26	5.29
30.28	2.95	1.82
32.57	2.75	1.07
34.69	2.59	2.33
35.80	2.51	2.02
36.56	2.46	1.51

Preparation of Form CS11 (Method 2)

The process of preparing Form CS11 comprises the following steps:

Dissoving step: about 10 mg of apabetalone solid was dissolved in corresponding solvent of Table 8.3, and then filtered to get clear solutions.

- 5 Precipitation step: The prepared solution was left at room temperature for slow evaporation until soild precipitated. The obtained solid was Form CS11 of Apabetalone.

Said reaction conditions, solvent composition and solvent amount of the preparation method of apabetalone Form CS11 are shown in Table 8.3. The samples 8-c to 8-f were confirmed to be

Form CS11 by XRPD. The XRPD pattern of sample 8-d is substantially as depicted in Figure 8D, and the XRPD data are listed in Table 8.4.

Table 8.3

Sample No	Weight (mg)	Solvent (v/v)	Volume (mL)	T (°C)
8-c	10.4	Chloroform	2.0	25
8-d	10.3	Methanol/acetone (1:1)	1.7	25
8-e	10.1	Methanol/methyl isobutyl ketone (2:1)	1.7	25
8-f	10.2	Methanol/toluene (4:1)	1.3	25

5

Table 8.4

2 Theta	d spacing	Intensity%
4.36	20.26	9.4
7.60	11.64	29.25
7.79	11.34	74.56
8.78	10.07	28.33
9.68	9.14	100
10.13	8.73	8.22
13.00	6.81	6.84
13.54	6.54	13.96
15.29	5.80	8.47
15.68	5.65	5.67
16.83	5.27	7.76
17.65	5.02	4.52
19.48	4.56	3.39
20.37	4.36	4.97
20.97	4.24	3.31
21.60	4.11	8.44
22.65	3.93	9.42
23.08	3.85	2.33
23.67	3.76	3.29
25.65	3.47	0.78
27.34	3.26	2.23
30.18	2.96	1.12

32.46	2.76	0.27
34.67	2.59	0.62
36.15	2.48	0.32

Stability study of Form CS11

Two samples of apabetalone Form CS11 were placed in constant temperature and humidity chambers at 25 °C/60% RH and 40 °C/75% RH for 6 weeks in open dishes. Crystalline form of the sample were tested by XRPD and impurity of the sample were tested to check the stability of Form CS11. The XRPD pattern overlay was substantially as depicted in Figure 8E (from top to bottom: XRPD pattern of Form CS11 before and after being stored under 25 °C/60% RH and 40 °C/75% RH for 6 weeks).

No form change and obvious purity decrease was observed for Form CS11 after being stored at 25 °C/60% RH and 40 °C/75% RH for 6 weeks. It can be seen that Form CS11 has good stability.

Solubility study of Form CS11

The prepared solid of apabetalone Form CS11 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, 4 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 8.5.

Table 8.5

Solubility (mg/mL) Time (h)	SGF	FeSSIF
1	0.71	0.29
4	0.65	0.26
24	0.65	0.30

The above results show that Form CS11 of apabetalone has good solubility in SGF and FeSSIF. Polymorph with high solubility is beneficial to increase the blood concentration of drugs in human body and improve the bioavailability of drugs, which is of great significance for drug research.

Example 9

Preparation of Form CS4

Form CS4 was obtained by heating Form CS11 of apabetalone to 220 °C.

The XRPD pattern of Form CS4 is substantially as depicted in Figure 9A, and the XRPD data are listed in Table 9.1. The DSC curve is substantially as depicted in Figure 9B, which shows the first

endothermic peak at around 231 °C. The TGA curve is substantially as depicted in Figure 9C, which shows about 1.1% weight loss when heated to 220 °C.

Table 9.1

2 Theta	d spacing	Intensity%
8.48	10.43	6.12
9.11	9.71	53.67
10.29	8.60	26.02
10.81	8.18	10.09
11.64	7.60	14.92
14.17	6.25	7.64
14.45	6.13	45.98
14.84	5.97	7.80
15.34	5.78	2.53
16.38	5.41	4.37
17.06	5.20	2.02
17.72	5.00	1.91
18.32	4.84	2.30
19.25	4.61	6.41
19.53	4.55	12.06
19.80	4.48	2.96
20.73	4.29	2.28
22.60	3.93	2.69
23.51	3.78	83.35
24.23	3.67	100.00
24.99	3.56	19.67
25.47	3.50	7.99
25.89	3.44	4.39
26.32	3.39	35.61
26.97	3.31	6.81
28.30	3.15	5.10
29.95	2.98	2.52
32.05	2.79	5.67
33.73	2.66	2.67
36.61	2.45	3.15
37.68	2.39	1.06

Long-term stability study of Form CS4

Two samples of apabetalone Form CS4 were placed in open dishes in constant temperature and humidity chambers at 25 °C/60% RH and 40 °C/75% RH for 10 months. Crystalline form of the sample were tested by XRPD and impurity of the sample were tested to check the stability of Form CS4. The stability results are substantially as depicted in Figure 9D (from top to bottom: XRPD pattern of Form CS4 before and after being stored under 25 °C/60% RH and 40 °C/75% RH for 10 months) and Table 9.2.

No form change and obvious purity decrease was observed for Form CS4 after being stored at 25 °C/60% RH and 40 °C/75% RH for 10 months. It can be seen that Form CS4 has good stability.

Table 9.2

Condition	1 week	2 weeks	4 weeks	10 months
25 °C /60%RH	99.10	99.06	99.05	99.04
40 °C /75%RH	99.10	98.99	99.04	99.02

Solubility study of Form CS4

The prepared solid of apabetalone Form CS4 was suspended into SGF (simulated gastric fluids) and FeSSIF (fed state simulated intestinal fluids, pH=5.0) to obtain saturated solutions. After being equilibrated for 1 h, 4 h and 24 h, concentrations of the saturated solutions were measured by HPLC. The results are listed in Table 9.3, which indicated that Form CS4 has good solubility.

Table 9.3

Solubility (mg/mL) Time (h)	SGF	FeSSIF
1	0.30	0.14
4	0.28	0.10
24	0.38	0.16

Hygroscopicity study of Form CS4

Dynamic vapor sorption (DVS) was applied to test hygroscopicity of Form CS4 of the present disclosure with about 10 mg of sample. The result was listed in Table 9.4. The DVS plot of Form CS4 is as depicted in Figure 9E.

Table 9.4

Form	Weight gain under 80% relative humidity
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Form CS4	0.12%
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The results showed that weight gain of Form CS4 under 80%RH is 0.12%. According to the hygroscopicity criteria, Form CS4 is almost non hygroscopic. The XRPD pattern of Form CS4 after DVS test is shown in Figure 9F. No form change was observed for Form CS4 before and after DVS test, which indicates that Form CS4 is stable under the influence of humidity.

Form CS4 of the present disclosure shows low hygroscopicity, which can well avoid issues such as crystal instability in the process of drug preparation and/or storage, as well as the unprocessability of the preparation caused by external factors such as environmental moisture, which is conducive to the accurate quantitative preparation and later transportation and storage.

Mechanical stability of Form CS4

Certain amount of Form CS4 was placed in a mortar and ground manually for 5 minutes. Crystalline form of the sample was checked by XRPD. The results are shown in Figure 9G.

The results showed that no form change and obvious crystalline decrease was observed for Form CS4 of apabetalone under certain mechanical stress. Form CS4 can maintain stable physical and chemical properties under certain mechanical stress, which is suitable for drug preparation and storage. Grinding of API is usually needed in the process of formulation, and good mechanical stability will reduce the risk of crystallinity decreasing and transformation of solid form of API.

The examples described above are only for illustrating the technical concepts and features of the present disclosure, and intended to make those skilled in the art being able to understand the present disclosure and thereby implement it, and should not be concluded to limit the protective scope of the present disclosure. Any equivalent variations or modifications according to the spirit of the present disclosure should be covered by the protective scope of the present disclosure.

THE CLAIMS DEFINING THE INVENTION ARE AS FOLLOWS:

1. A crystalline form CS1 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $6.1^{\circ} \pm 0.2^{\circ}$, $12.3^{\circ} \pm 0.2^{\circ}$, $13.0^{\circ} \pm 0.2^{\circ}$, $14.1^{\circ} \pm 0.2^{\circ}$, $16.4^{\circ} \pm 0.2^{\circ}$, $17.1^{\circ} \pm 0.2^{\circ}$, $18.5^{\circ} \pm 0.2^{\circ}$, $20.5^{\circ} \pm 0.2^{\circ}$, $23.2^{\circ} \pm 0.2^{\circ}$, $24.5^{\circ} \pm 0.2^{\circ}$, $26.1^{\circ} \pm 0.2^{\circ}$ and $26.8^{\circ} \pm 0.2^{\circ}$ using $\text{CuK}\alpha$ radiation.

2. A process for preparing crystalline form CS1 according to claim 1, the process comprising method 1) or method 2), wherein:

method 1) comprises:

a dissolving step of dissolving apabetalone solid in a solvent to obtain a solution; and

a precipitation step of adding the solution prepared in the dissolving step to an anti-solvent, or adding an anti-solvent to the solution prepared in the dissolving step, to obtain apabetalone Form CS1 solid;

dimethyl sulfoxide and dimethyl acetamide; and wherein said anti-solvent comprises one or two solvents selected from n-heptane, methyl tert-butyl ether, toluene, water and acetonitrile; and

method 2) comprises:

a dissolving step of dissolving apabetalone solid in a solvent at 40°C to 60°C to obtain a solution; and

a precipitation step of cooling the solution prepared in the dissolving step to -20°C to 5°C to obtain apabetalone Form CS1 solid;

wherein said solvent comprises tetrahydrofuran, acetone, a mixture of tetrahydrofuran and methyl tert-butyl ether, a mixture of ethyl acetate and acetone, or a mixture of acetonitrile and N,N-dimethyl formamide.

3. The process according to claim 2, wherein in method 2), the solvent comprises: tetrahydrofuran and methyl tert-butyl ether at a volume ratio of 2:1; thyl acetate and acetone at a volume ratio of 1:1; or acetonitrile and N,N-dimethyl formamide at a volume ratio of 9:1.

4. A crystalline form CS2 of apabetalone, wherein the X-ray powder diffraction pattern

shows characteristic peaks at 2theta values of $25.3^{\circ} \pm 0.2^{\circ}$, $23.1^{\circ} \pm 0.2^{\circ}$, $20.2^{\circ} \pm 0.2^{\circ}$, $15.3^{\circ} \pm 0.2^{\circ}$, $13.3^{\circ} \pm 0.2^{\circ}$, $11.5^{\circ} \pm 0.2^{\circ}$, $6.6^{\circ} \pm 0.2^{\circ}$, $5.1^{\circ} \pm 0.2^{\circ}$, and $8.8^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

5 5. A process for preparing crystalline form CS2 according to claim 4, wherein the process comprises: forming a suspension by adding apabetalone solid into one or more alcohols, a mixture of alcohols and ethers, a mixture of alcohols and ketones, or a mixture of alcohols and arenes; heating the suspension to dissolve the apabetalone solid, and then filtering and cooling the resulting solution to obtain precipitation; wherein the solid obtained by the precipitation comprises Form CS2 of apabetalone.

10 6. The process according to claim 5, wherein said alcohol comprises methanol; said ether comprises 2-methyltetrahydrofuran; said arene comprises toluene; said ketone comprises methyl isobutyl ketone; said heating temperature is 50°C to 100°C; and said cooling temperature is -20°C to 5°C.

15 7. A crystalline form CS8 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $28.1^{\circ} \pm 0.2^{\circ}$, $25.9^{\circ} \pm 0.2^{\circ}$, $23.9^{\circ} \pm 0.2^{\circ}$, $20.2^{\circ} \pm 0.2^{\circ}$, $13.5^{\circ} \pm 0.2^{\circ}$, $13.1^{\circ} \pm 0.2^{\circ}$, $9.7^{\circ} \pm 0.2^{\circ}$, $7.8^{\circ} \pm 0.2^{\circ}$, $22.5^{\circ} \pm 0.2^{\circ}$ and $11.4^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

20 8. A process for preparing crystalline form CS8 according to ae of claim 7, wherein the process comprises: forming a suspension by adding apabetalone solid into a single solvent comprising halohydrocarbons, or a mixture of solvents comprising halohydrocarbons and alcohols; heating the suspension to dissolve the apabetalone solid, and then filtering and cooling the resulting solution to obtain precipitation; wherein the solid obtained from the precipitation comprises Form CS8 of apabetalone.

25 9. The process according to claim 8, wherein, said halohydrocarbon comprises dichloromethane; said alcohol comprises isopropanol; said volume ratio of halohydrocarbon and isopropanol is 4:1; said heating temperature comprises 40°C to 60°C, said cooling temperature comprises -20°C to 5°C.

10. A crystalline form CS13 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $5.1^{\circ} \pm 0.2^{\circ}$, $6.4^{\circ} \pm 0.2^{\circ}$, $7.8^{\circ} \pm 0.2^{\circ}$, $8.5^{\circ} \pm 0.2^{\circ}$, $12.5^{\circ} \pm 0.2^{\circ}$, $16.0^{\circ} \pm 0.2^{\circ}$, $17.1^{\circ} \pm 0.2^{\circ}$ and $25.7^{\circ} \pm 0.2^{\circ}$, using CuK α radiation.
- 5 11. A process for preparing crystalline form CS13 according to claim 10, wherein the process comprises: forming a solution by dissolving apabetalone solid in a mixture of ethers and water or a mixture of ketones and water; filtering the solution and then evaporating the filtrate at room temperature to obtain a solid; wherein the obtained solid comprises Form CS13 of apabetalone.
- 10 12. The process according to claim 11, wherein said ether comprises tetrahydrofuran; said ketone comprises acetone; said volume ratio of ether and water is 4:1; said volume ratio of ketone and water is 9:1.
- 15 13. A crystalline form CS20 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $5.6^{\circ} \pm 0.2^{\circ}$, $8.4^{\circ} \pm 0.2^{\circ}$, $9.4^{\circ} \pm 0.2^{\circ}$, $11.3^{\circ} \pm 0.2^{\circ}$, $13.5^{\circ} \pm 0.2^{\circ}$, $14.5^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$, $18.9^{\circ} \pm 0.2^{\circ}$, $20.1^{\circ} \pm 0.2^{\circ}$, $20.6^{\circ} \pm 0.2^{\circ}$, $22.8^{\circ} \pm 0.2^{\circ}$, $24.4^{\circ} \pm 0.2^{\circ}$ and $26.3^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.
- 20 14. A process for preparing crystalline form CS20 according to claim 13, wherein the process comprises: forming a solution by dissolving apabetalone solid in a mixture of solvents selected from acetic acid and nitriles or acetic acid and esters; filtering the solution and then evaporating the filtrate at room temperature to obtain a solid; wherein the obtained solid comprises Form CS20 of apabetalone.
- 25 15. The process according to claim 14, wherein, said nitrile comprises acetonitrile; said ester comprises ethyl acetate; said volume ratio of nitrile and acetic acid is 9:1; said volume ratio of ester and acetic acid is 4:1.
- 30 16. A crystalline form CS7 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $5.9^{\circ} \pm 0.2^{\circ}$, $6.7^{\circ} \pm 0.2^{\circ}$, $8.4^{\circ} \pm 0.2^{\circ}$, $10.7^{\circ} \pm 0.2^{\circ}$, $12.5^{\circ} \pm 0.2^{\circ}$, $13.3^{\circ} \pm 0.2^{\circ}$, $15.0^{\circ} \pm 0.2^{\circ}$, $16.0^{\circ} \pm 0.2^{\circ}$, $16.9^{\circ} \pm 0.2^{\circ}$, $21.8^{\circ} \pm 0.2^{\circ}$, $24.5^{\circ} \pm 0.2^{\circ}$ and $25.1^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

17. A process for preparing crystalline form CS7 according to claim 16, wherein the process comprises:

5 a dissolving step of dissolving apabetalone solid in a single solvent of halohydrocarbons to obtain a clear solution;

a precipitation step of inducing the formation of a solid by liquid vapor diffusion; wherein the solid obtained by the liquid vapor diffusion comprises Form CS7 of apabetalone;

10 or wherein said precipitation step comprises: storing the clear solution prepared in the dissolving step in an open glass vial; and then placing the open glass vial into another vial containing ketones until a solid is obtained; wherein the obtained solid comprises Form CS7 of apabetalone.

18. The process according to claim 17, wherein, said halohydrocarbon comprises chloroform; and said ketone comprises methyl isobutyl ketone.

19. A crystalline form CS9 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $3.9^{\circ} \pm 0.2^{\circ}$, $6.0^{\circ} \pm 0.2^{\circ}$, $7.3^{\circ} \pm 0.2^{\circ}$, $9.9^{\circ} \pm 0.2^{\circ}$, $12.1^{\circ} \pm 0.2^{\circ}$, $12.8^{\circ} \pm 0.2^{\circ}$, $13.4^{\circ} \pm 0.2^{\circ}$, $17.0^{\circ} \pm 0.2^{\circ}$, $22.5^{\circ} \pm 0.2^{\circ}$, $24.2^{\circ} \pm 0.2^{\circ}$ and $24.9^{\circ} \pm 0.2^{\circ}$ using CuK α radiation.

20. A process for preparing crystalline form CS9 according to claim 19, wherein the process comprises:

a dissolving step of dissolving apabetalone solid in a mixture of solvents selected ethers and alcohols or halohydrocarbons and alcohols to obtain a clear solution;

25 a precipitation step of evaporating the prepared clear solution at room temperature to obtain a solid; wherein the obtained solid comprises Form CS9 of apabetalone.

21. The process according to claim 20, wherein, said ether comprises tetrahydrofuran, said alcohol comprises isopropanol; said halohydrocarbons comprises dichloromethane and chloroform; said volume ratio of ether and alcohol is 1:1; said volume ratio of halohydrocarbon and alcohol is 4:1.

22. A crystalline form CS11 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $4.4^{\circ}\pm 0.2^{\circ}$, $7.6^{\circ}\pm 0.2^{\circ}$, $7.8^{\circ}\pm 0.2^{\circ}$, $8.8^{\circ}\pm 0.2^{\circ}$, $9.7^{\circ}\pm 0.2^{\circ}$, $13.0^{\circ}\pm 0.2^{\circ}$, $13.6^{\circ}\pm 0.2^{\circ}$, $15.3^{\circ}\pm 0.2^{\circ}$, $16.9^{\circ}\pm 0.2^{\circ}$, $17.6^{\circ}\pm 0.2^{\circ}$, $21.6^{\circ}\pm 0.2^{\circ}$ and $22.7^{\circ}\pm 0.2^{\circ}$ using CuK α radiation.

23. A process for preparing crystalline form CS11 according to claim 22, the process comprises method 1) or method 2), wherein:

method 1) comprises:

a dissolving step of dissolving apabetalone solid in alcohols to obtain a clear solution;

a precipitation step of adding the prepared clear solution to water or adding water to the prepared clear solution to obtain a solid; wherein the obtained solid comprises Form CS11 of apabetalone; and

method 2) comprises:

a dissolving step of dissolving apabetalone solid in a single solvent of halohydrocarbons or a mixture of solvents selected from alcohols and ketones or alcohols and arenes to obtain a clear solution;

a precipitation step of evaporating the prepared clear solution at room temperature to obtain apabetalone Form CS11 solid.

24. The process according to claim 23, wherein, said alcohol of method 1) comprises methanol, said alcohol of method 2) comprises methanol; said ketones comprise acetone and methyl isobutyl ketone; said mixture of alcohol and ketone has a volume ratio of 1:1 to 2:1; and said mixture of alcohol and arene has a volume ratio of 4:1.

25. A crystalline form CS4 of apabetalone, wherein the X-ray powder diffraction pattern shows characteristic peaks at 2theta values of $9.1^{\circ}\pm 0.2^{\circ}$, $10.3^{\circ}\pm 0.2^{\circ}$, $10.8^{\circ}\pm 0.2^{\circ}$, $11.6^{\circ}\pm 0.2^{\circ}$, $14.5^{\circ}\pm 0.2^{\circ}$, $19.5^{\circ}\pm 0.2^{\circ}$, $23.5^{\circ}\pm 0.2^{\circ}$, $24.2^{\circ}\pm 0.2^{\circ}$, $25.0^{\circ}\pm 0.2^{\circ}$ and $26.3^{\circ}\pm 0.2^{\circ}$, using CuK α radiation.

26. A process for preparing crystalline form CS4 according to claim 25, wherein the process comprises: heating Form CS11 of apabetalone to 200°C to 220°C, wherein the resulting solid obtained from the heating step comprises Form CS4 of apabetalone.

27. A pharmaceutical composition, wherein said pharmaceutical composition comprises a therapeutically effective amount of one or more crystalline forms selected from Form CS1 according to claim 1, Form CS2 according to claim 4, Form CS8 according to claim 7, Form CS13 according to claim 10, Form CS20 according to claim 13, Form CS7 according to claim 16, Form CS9 according to claim 19, Form CS11 according to claim 22 and Form CS4 according to claim 25, and at least one pharmaceutically acceptable carrier, diluent or excipient.

28. A composition comprising one or more crystalline forms selected from Form CS1 according to claim 1, Form CS2 according to claim 4, Form CS8 according to claim 7, Form CS13 according to claim 10, Form CS20 according to claim 13, Form CS7 according to claim 16, Form CS9 according to claim 19, Form CS11 according to claim 22 or Form CS4 according to claim 25, for use in preparing a drug for treating cardiovascular, cholesterol or lipid-related disorders.

29. A composition comprising Form CS1 according to claim 1, Form CS2 according to claim 4, Form CS8 according to claim 7, Form CS13 according to claim 10, Form CS20 according to claim 13, Form CS7 according to claim 16, Form CS9 according to claim 19, Form CS11 according to claim 22 or Form CS4 according to claim 25, for use in preparing a drug for treating atherosclerosis, acute coronary syndrome, or predecessor diabetes.

30. A method of treating cardiovascular, cholesterol or lipid-related disorders in a subject in need thereof, comprising administering a therapeutically effective amount of one or more crystalline forms selected from Form CS1 according to claim 1, Form CS2 according to claim 4, Form CS8 according to claim 7, Form CS13 according to claim 10, Form CS20 according to claim 13, Form CS7 according to claim 16, Form CS9 according to claim 19, Form CS11 according to claim 22 and Form CS4 according to claim 15, and pharmaceutically acceptable carriers, diluents or excipients.

31. A method of treating atherosclerosis and acute coronary syndrome and predecessor diabetes in a subject in need thereof, comprising administering a therapeutically effective amount of one or more crystalline forms selected from Form CS1 according to claim 1,

Form CS2 according to claim 4, Form CS8 according to claim 7, Form CS13 according to claim 10, Form CS20 according to claim 13, Form CS7 according to claim 16, Form CS9 according to claim 19, Form CS11 according to claim 22 and Form CS4 according to claim 25, and pharmaceutically acceptable carriers, diluents or excipients.

DRAWING

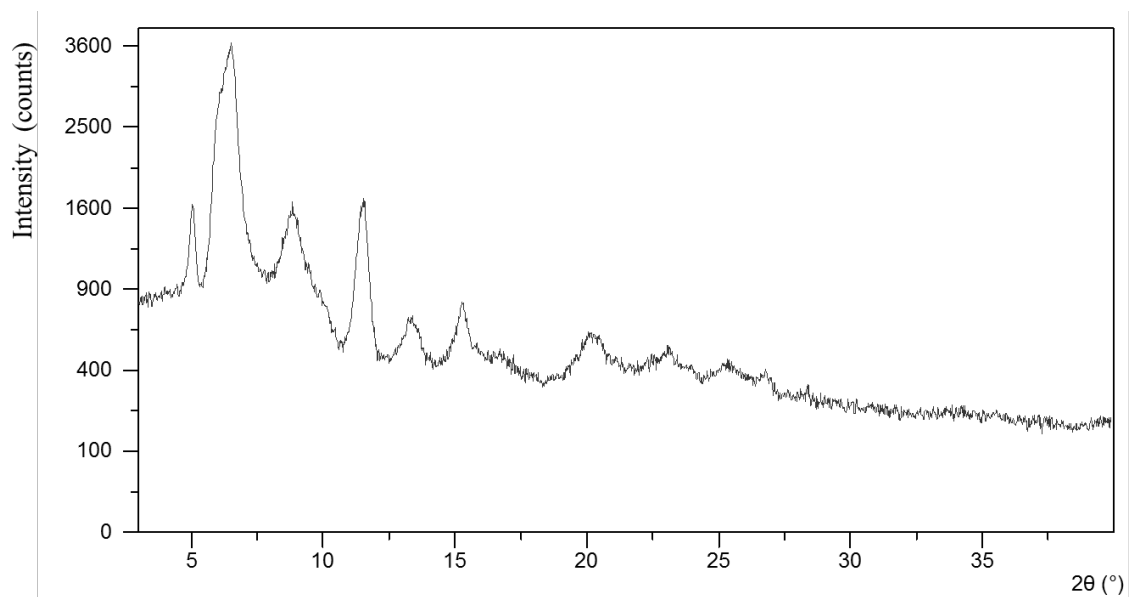


FIG. 1A

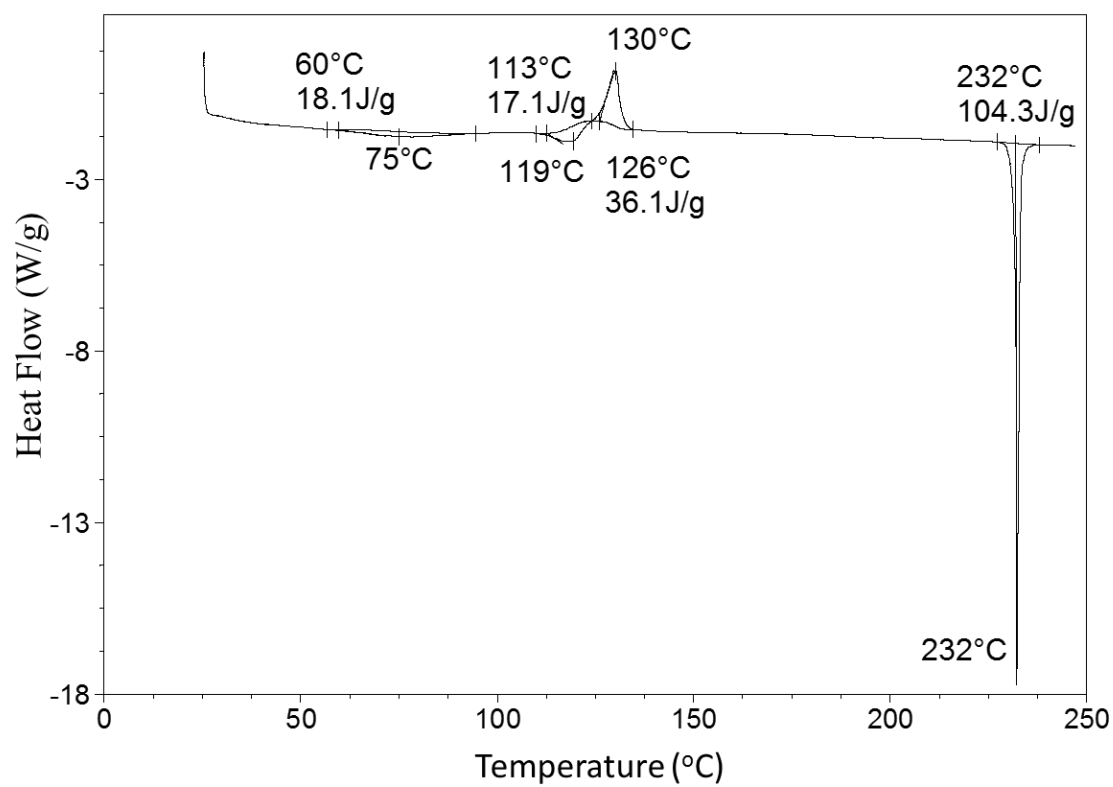


FIG. 1B

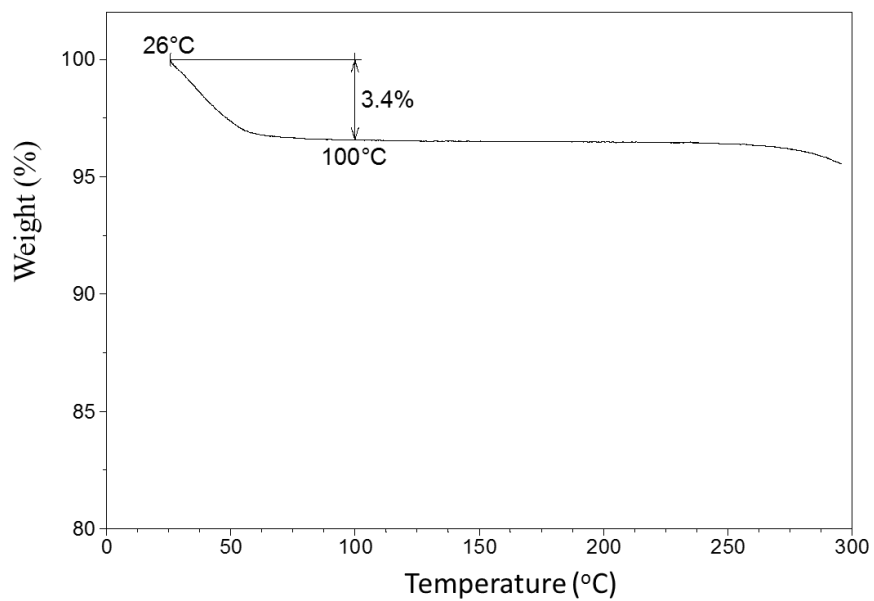


FIG. 1C

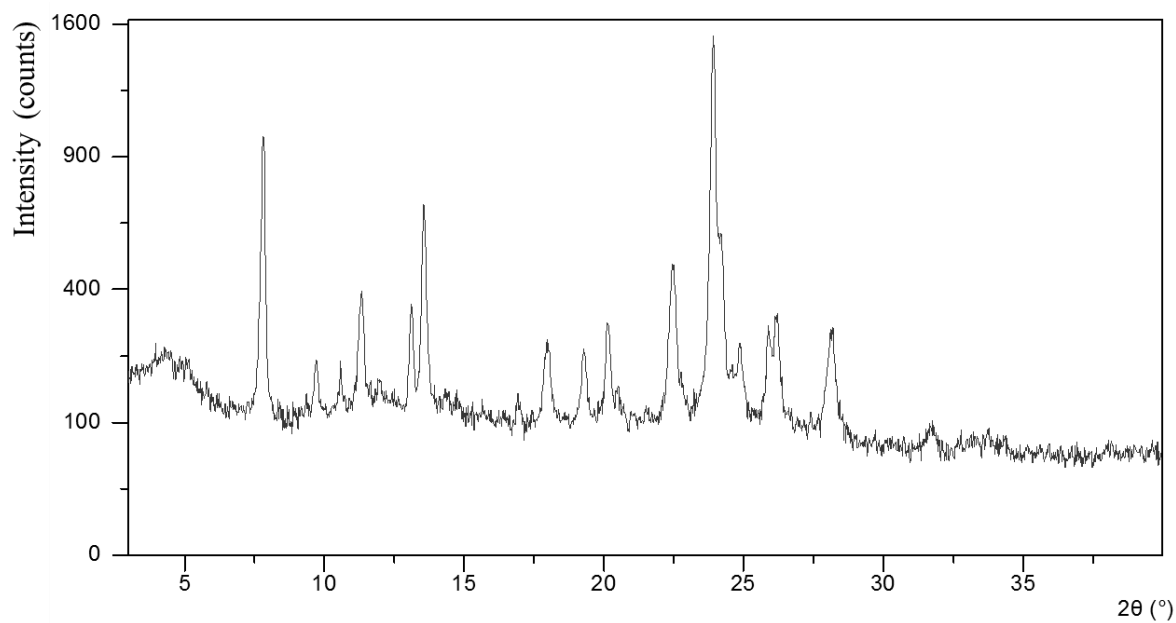


FIG. 2A

DRAWING

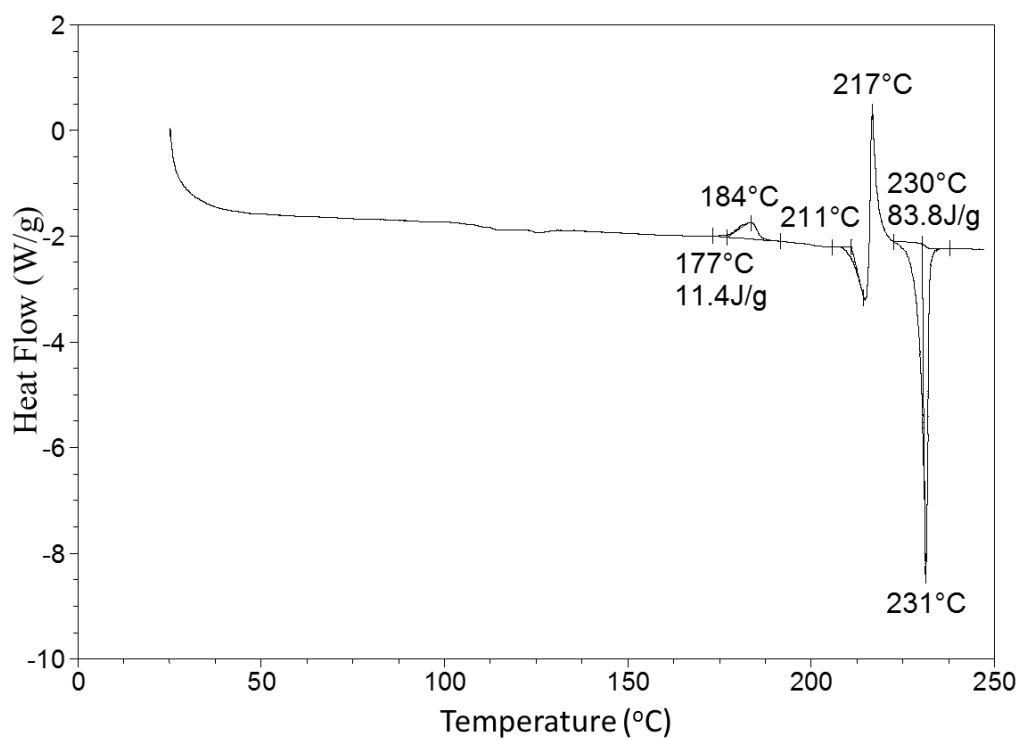


FIG. 2B

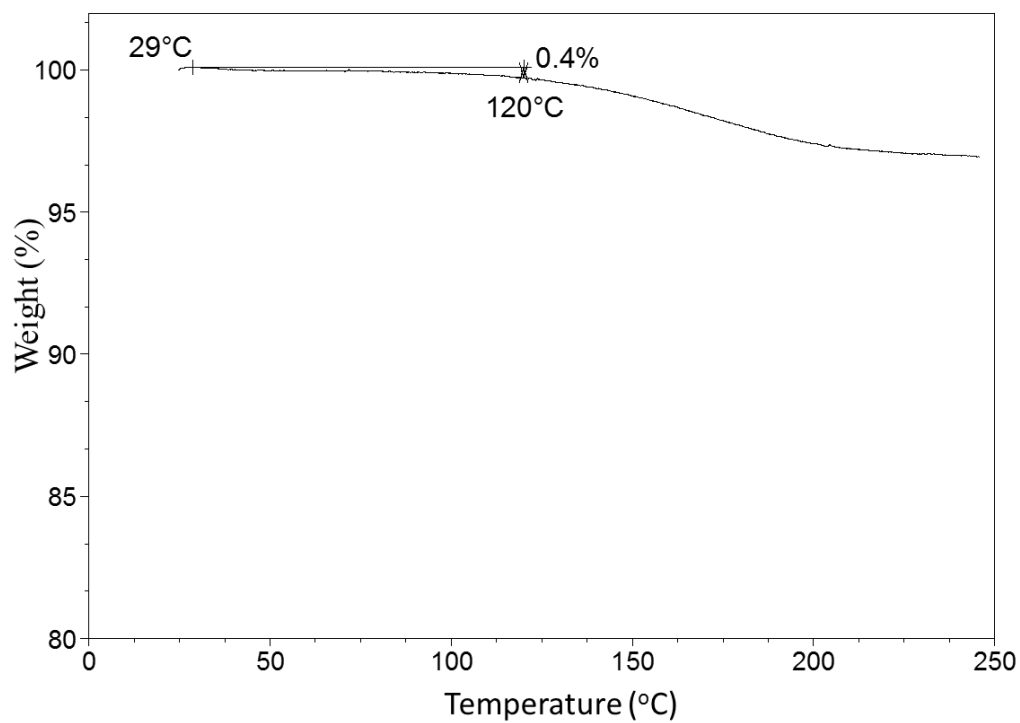


FIG. 2C

D R A W I N G

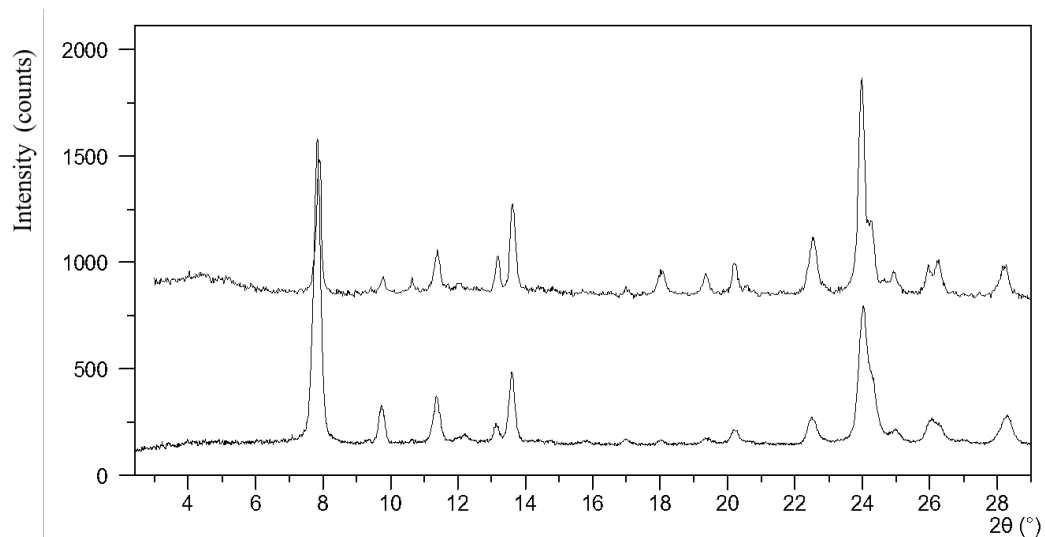


FIG. 2D

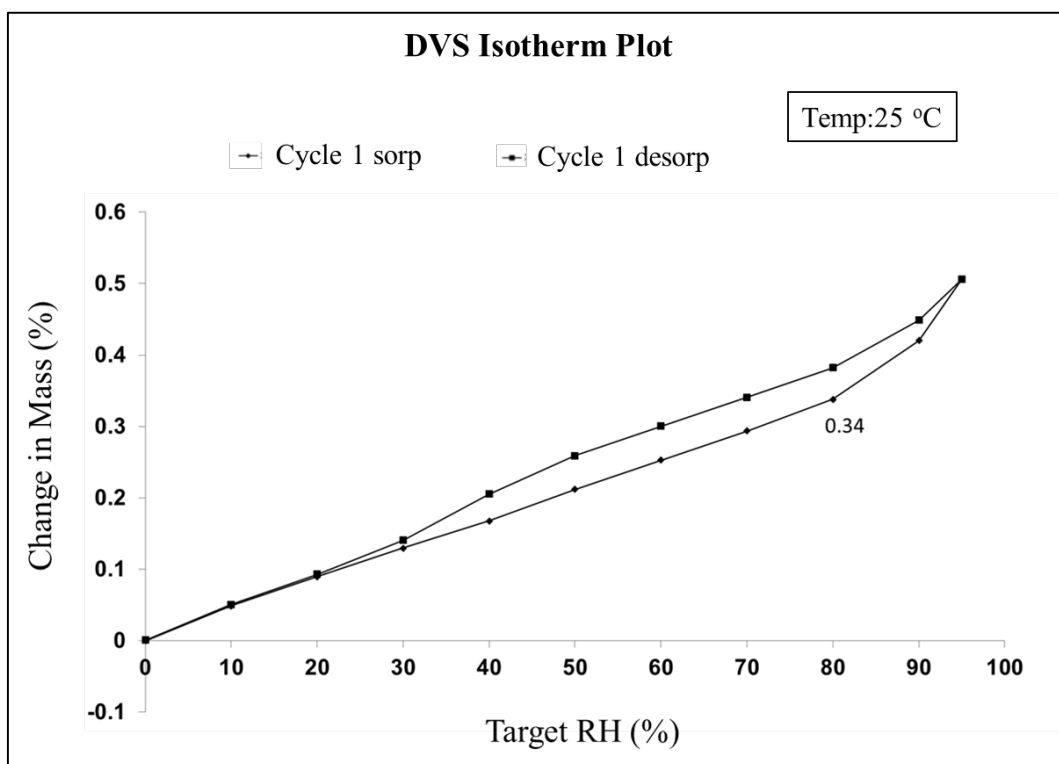


FIG. 2E

DRAWING

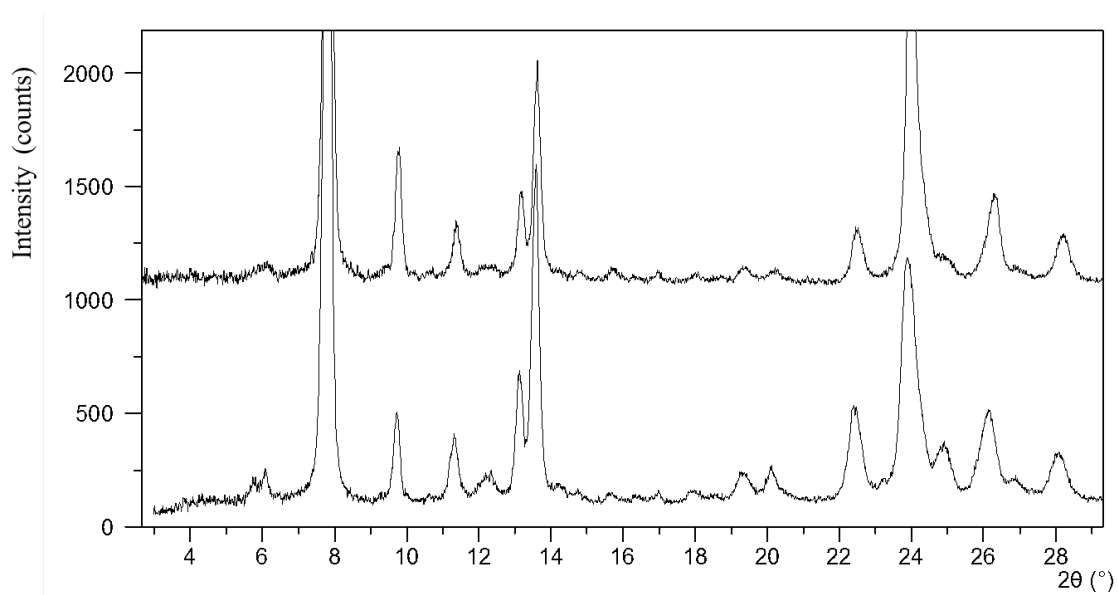


FIG. 2F

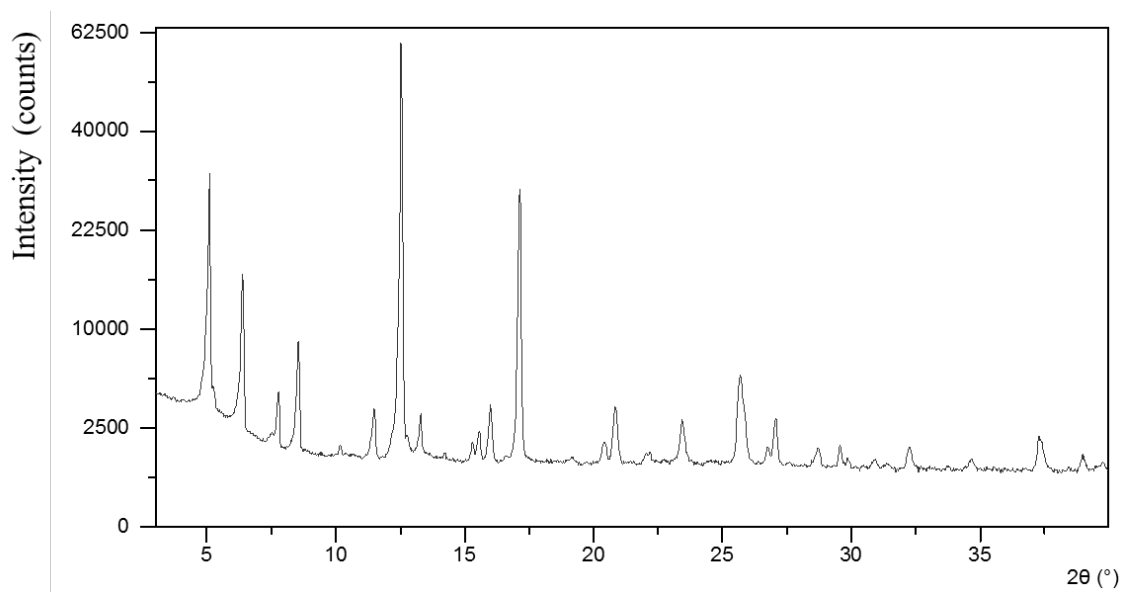


FIG. 3A

DRAWING

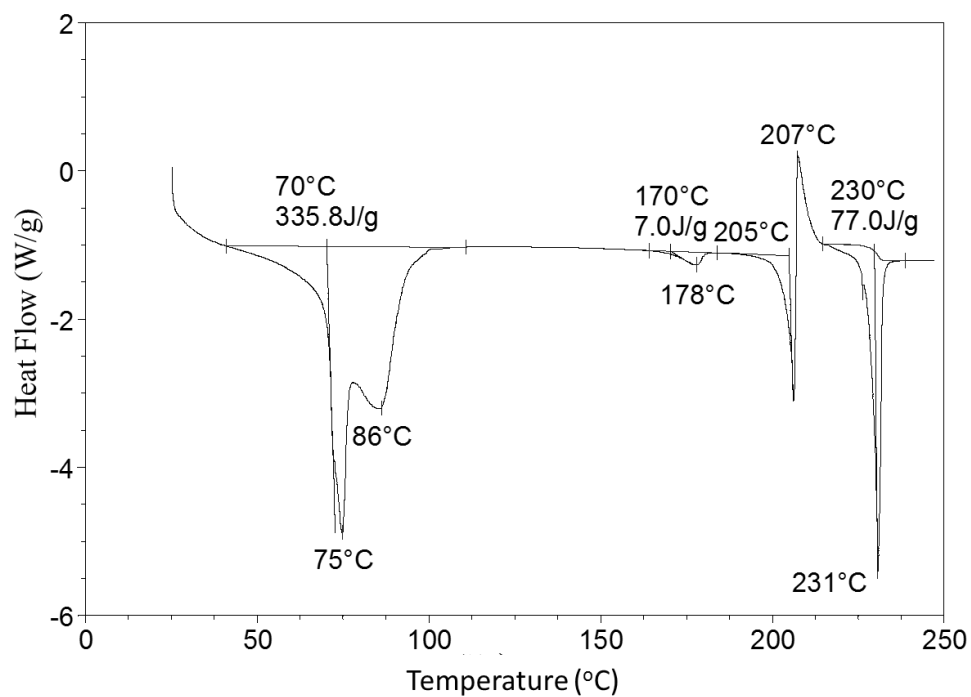


FIG. 3B

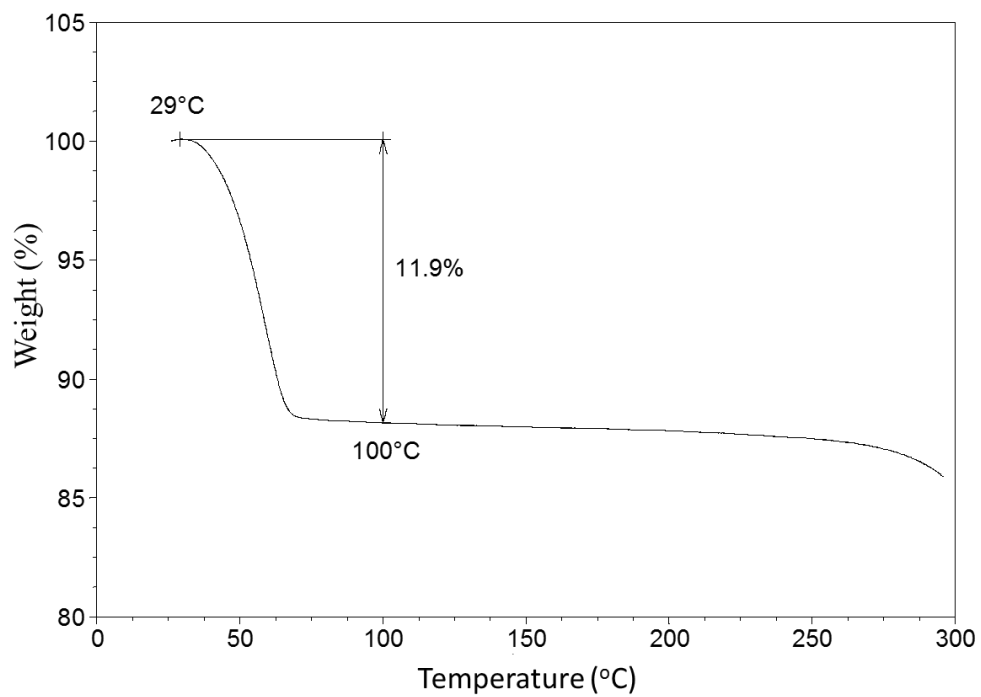


FIG. 3C

DRAWING

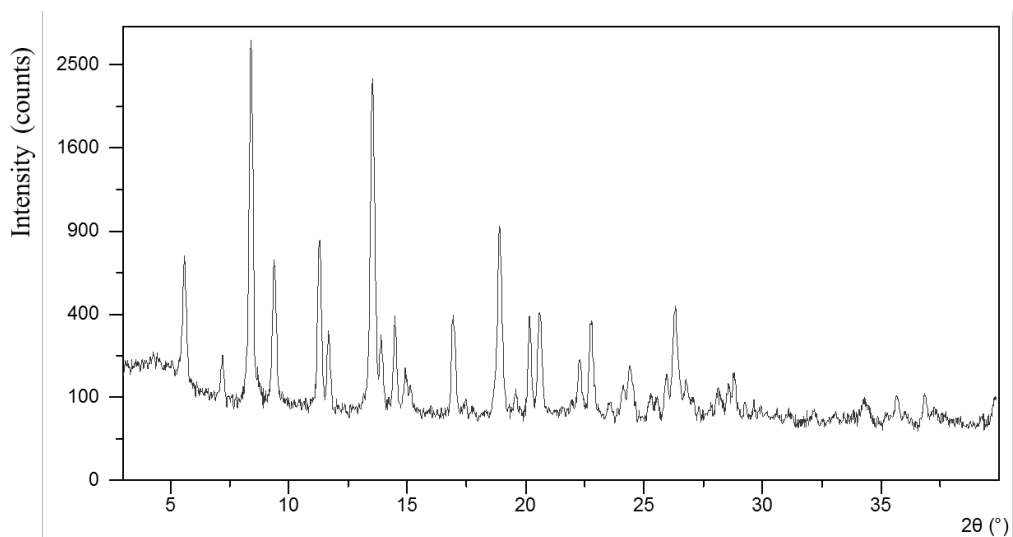


FIG. 4A

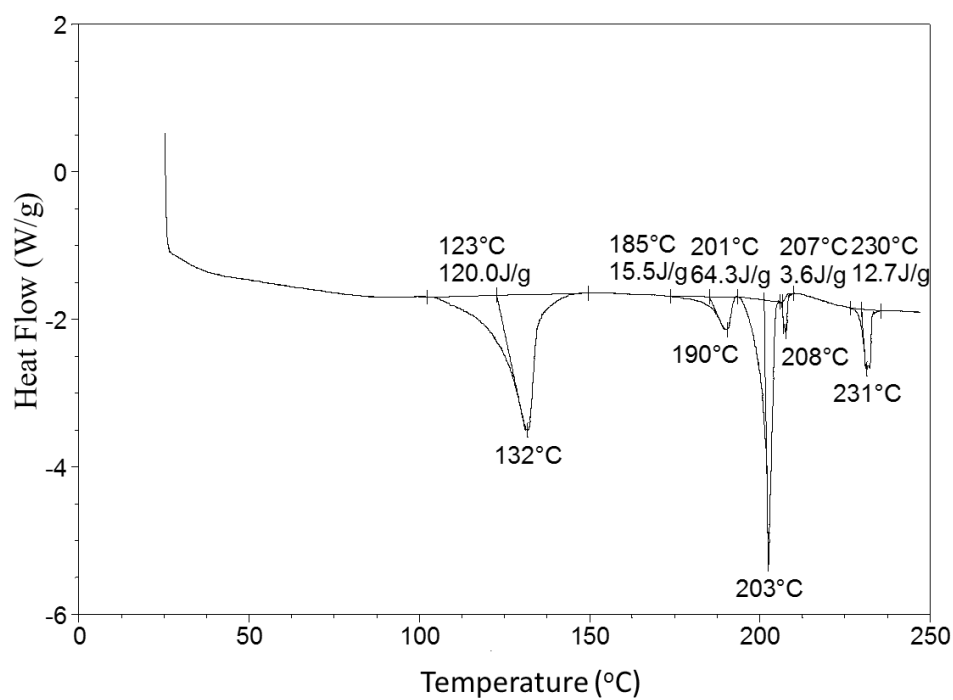


FIG. 4B

DRAWING

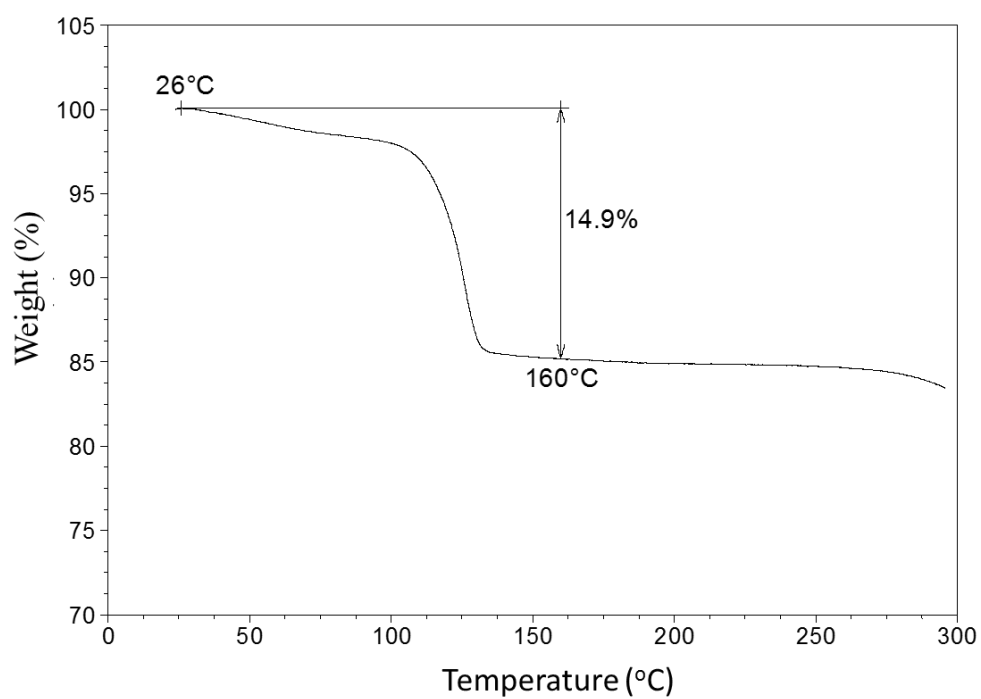


FIG. 4C

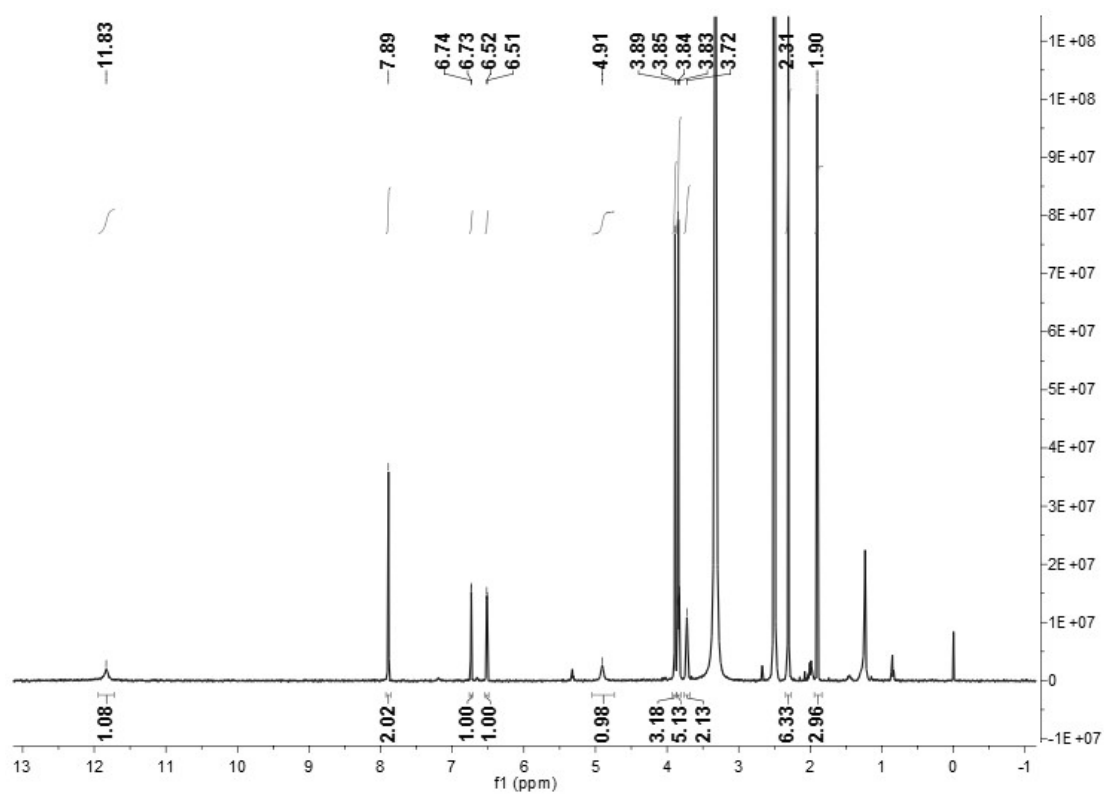


FIG. 4D

DRAWING

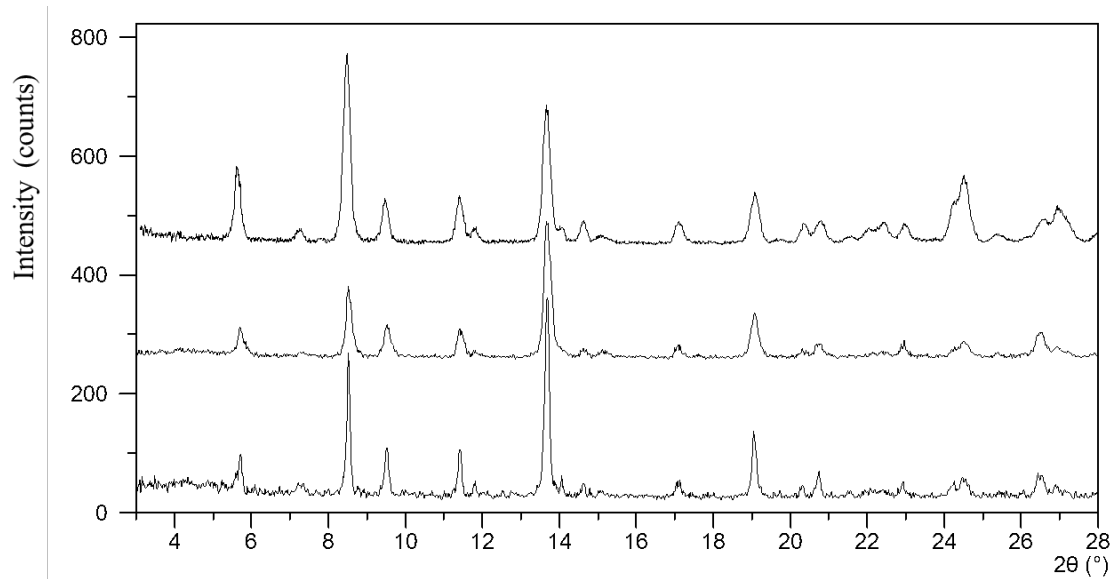


FIG. 4E

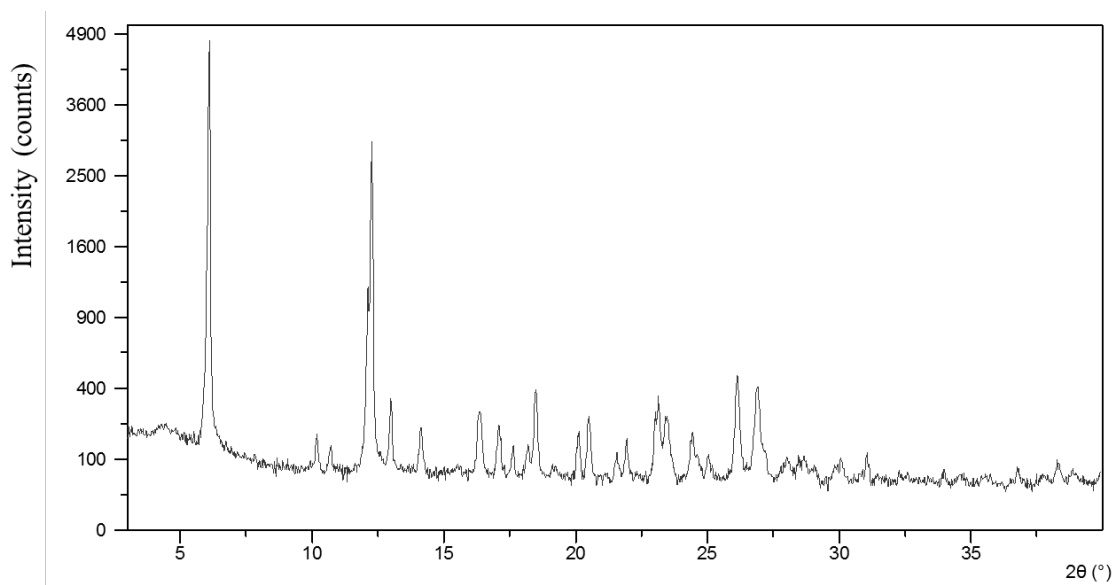


FIG. 5A

D R A W I N G

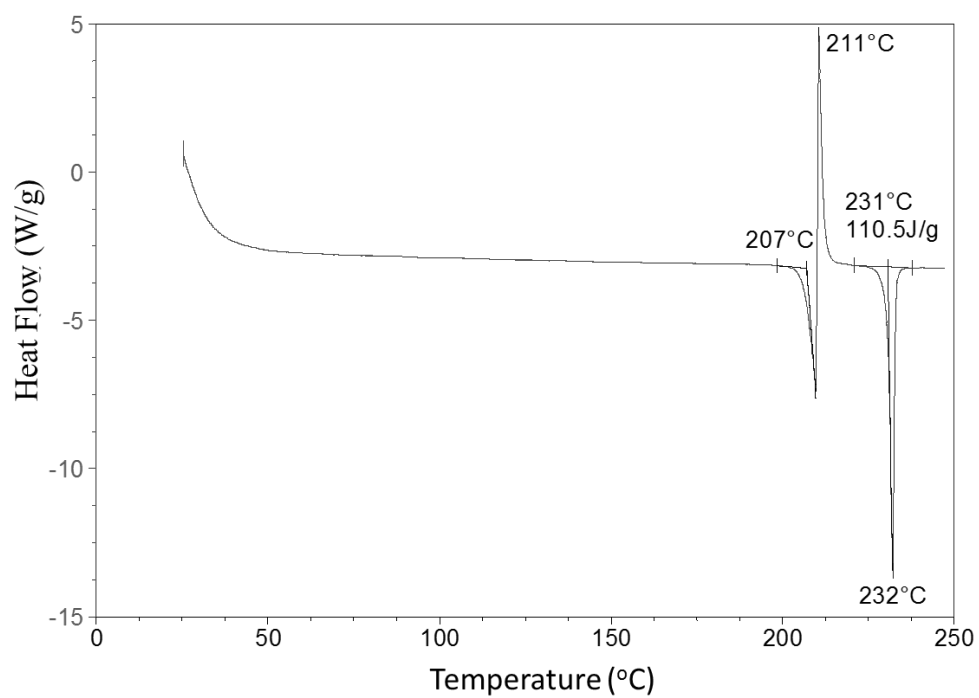


FIG. 5B

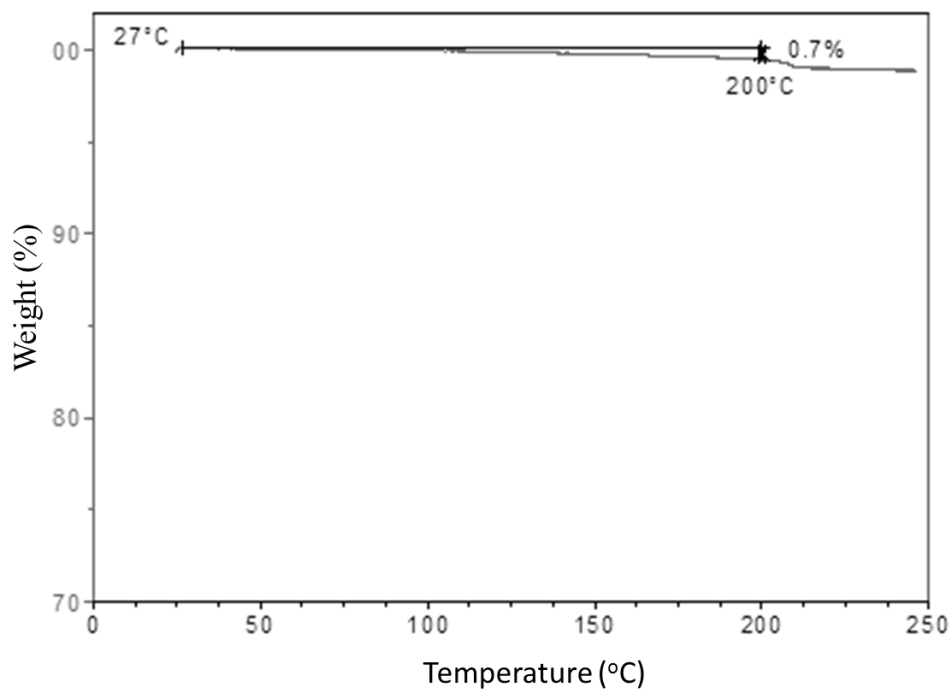


FIG. 5C

D R A W I N G

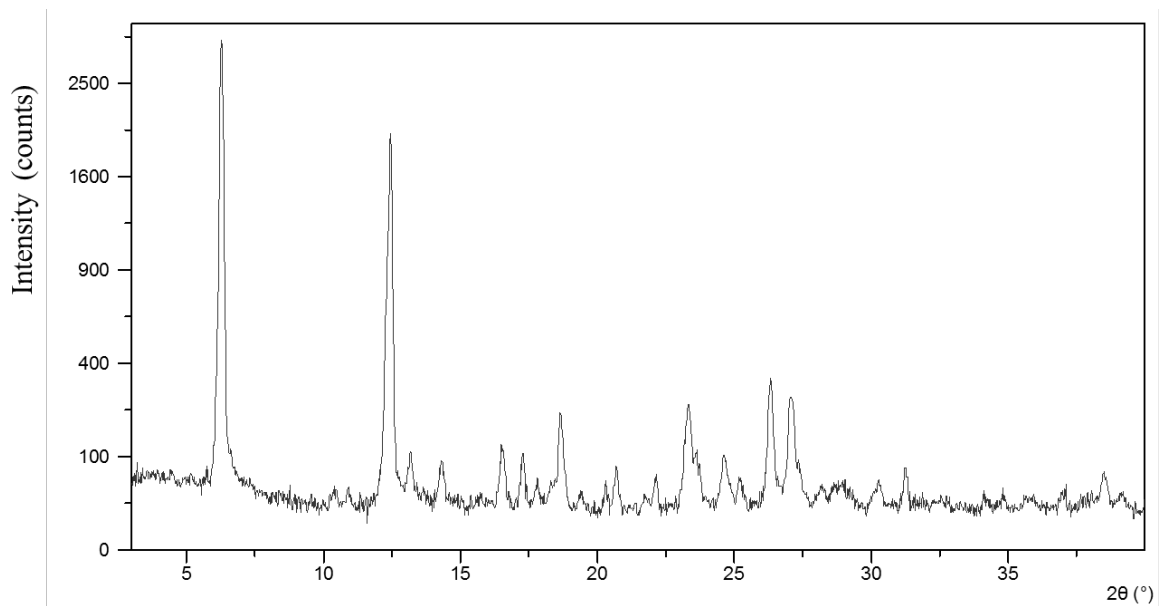


FIG. 5D

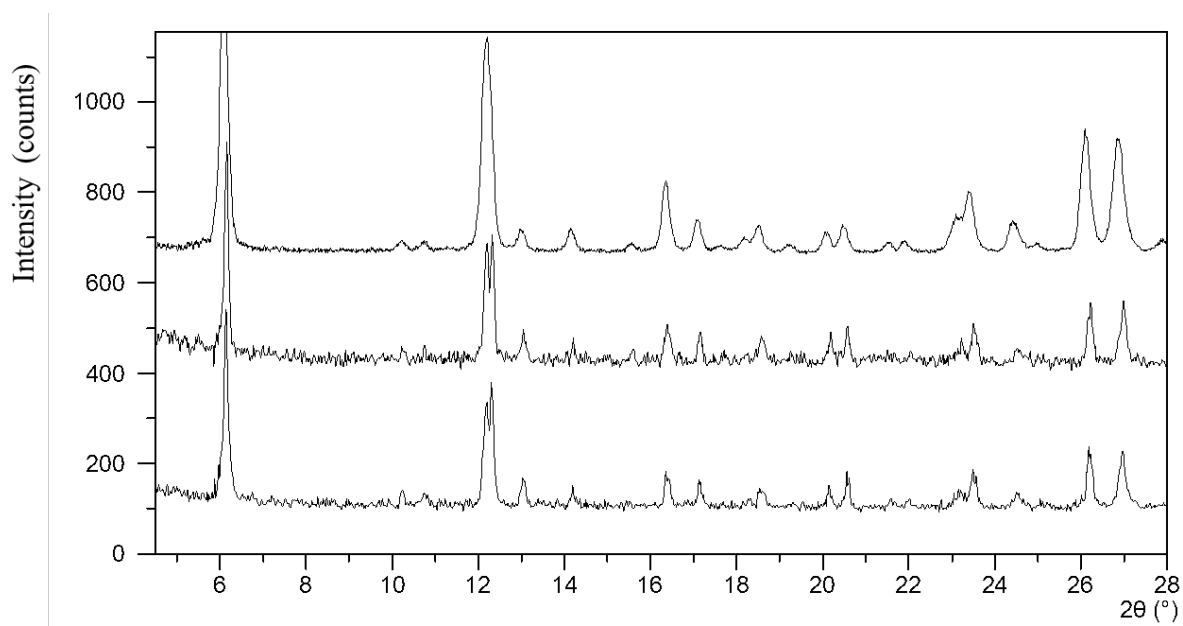


FIG. 5E

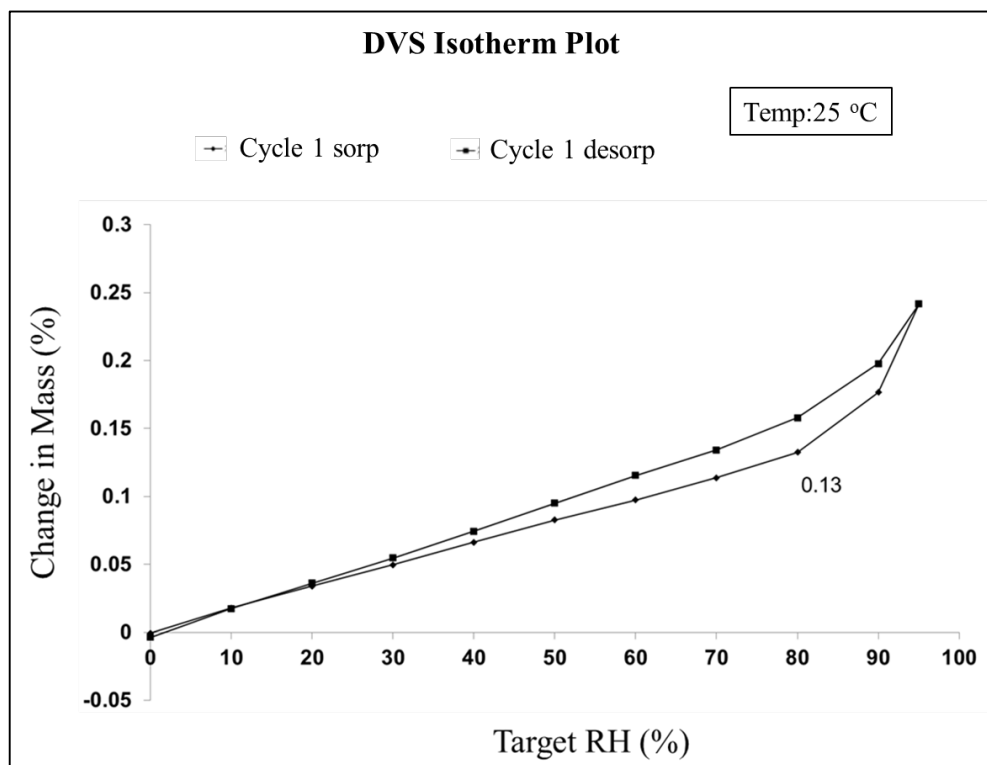


FIG. 5F

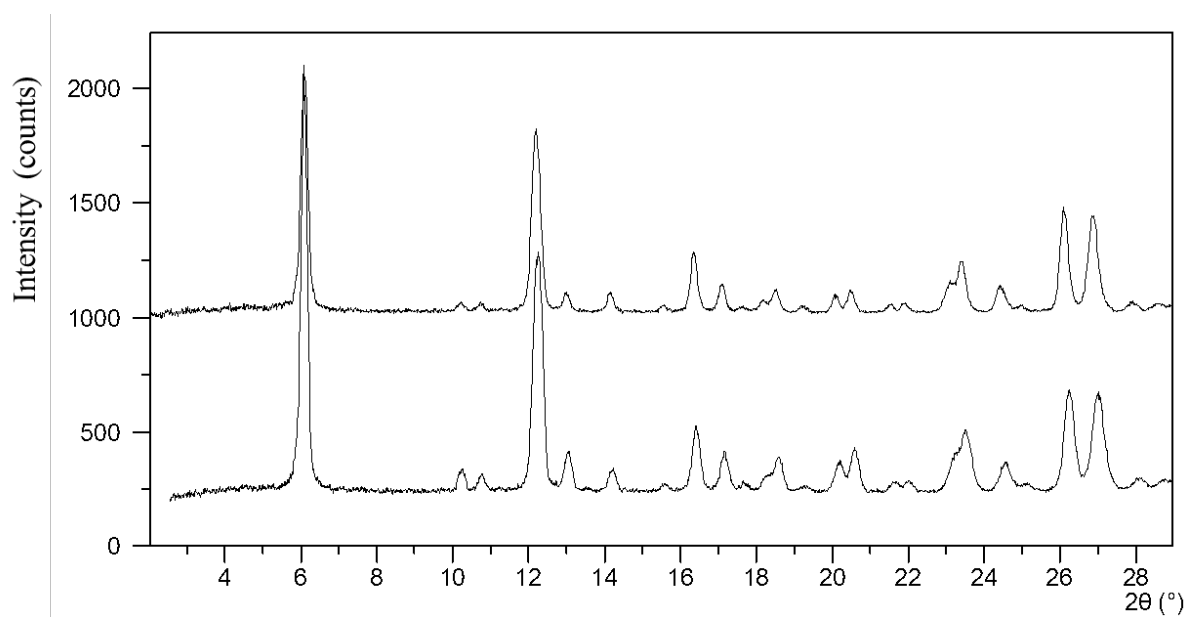


FIG. 5G

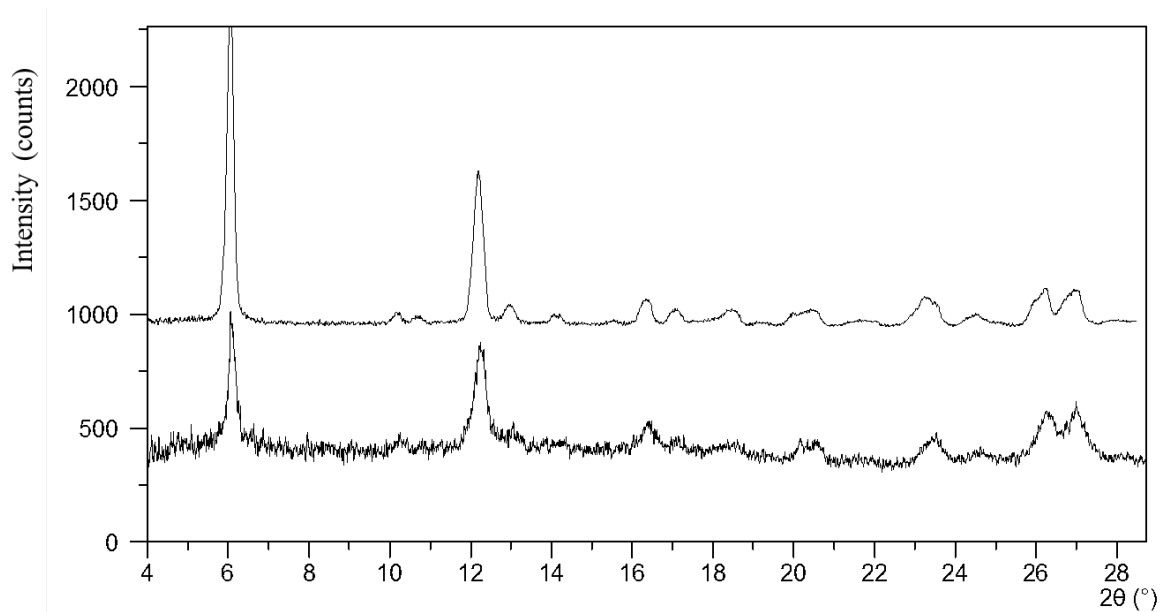


FIG. 5H

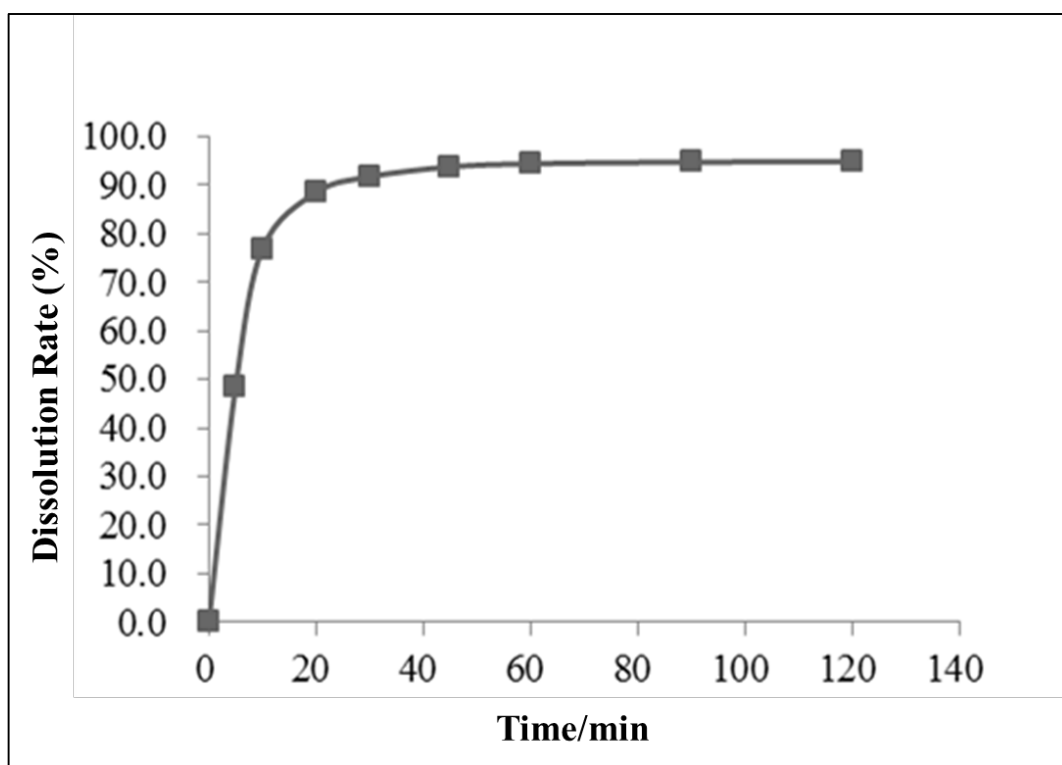


FIG. 5I

DRAWING

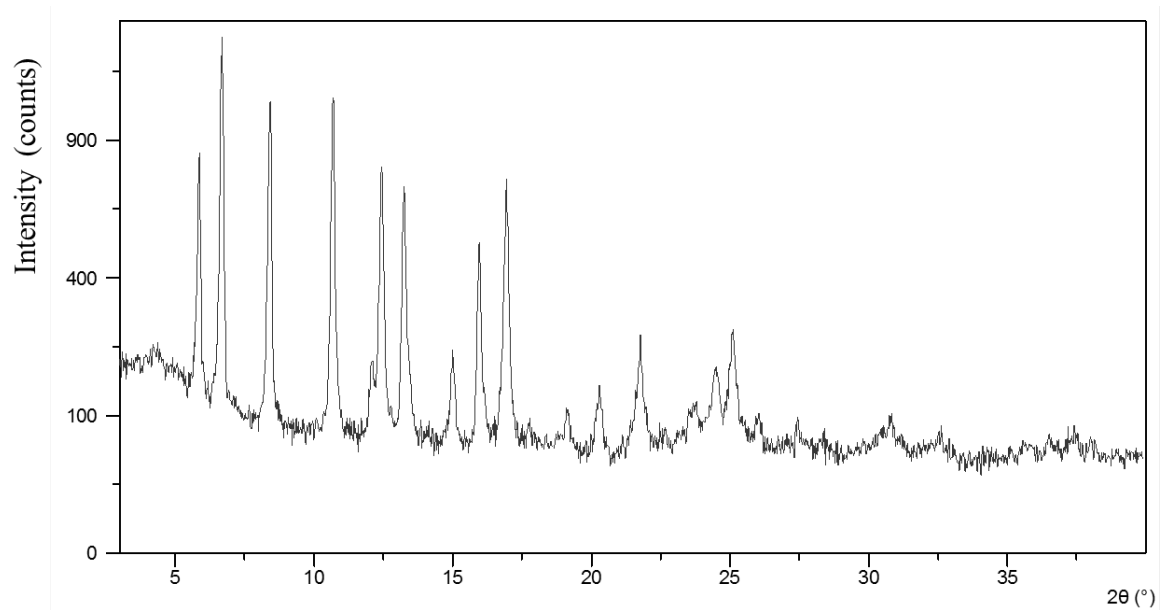


FIG. 6A

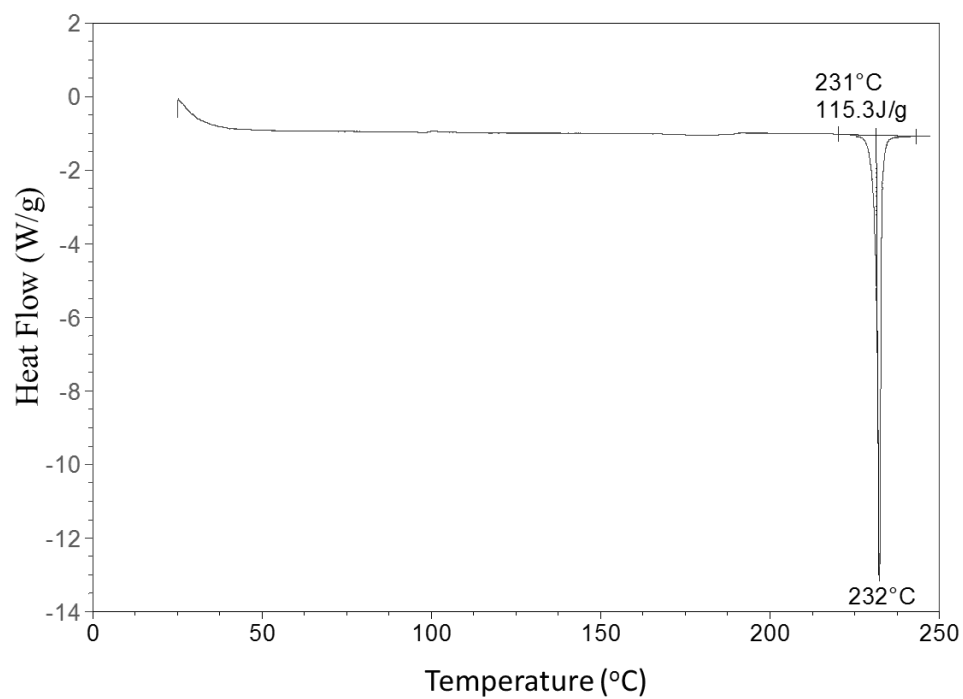


FIG. 6B

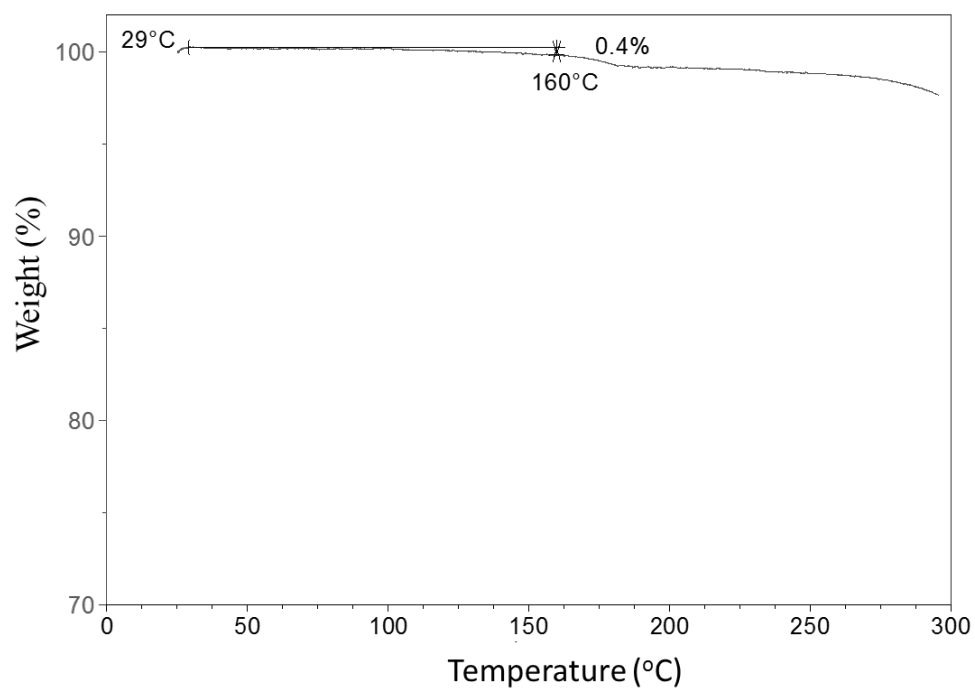


FIG. 6C

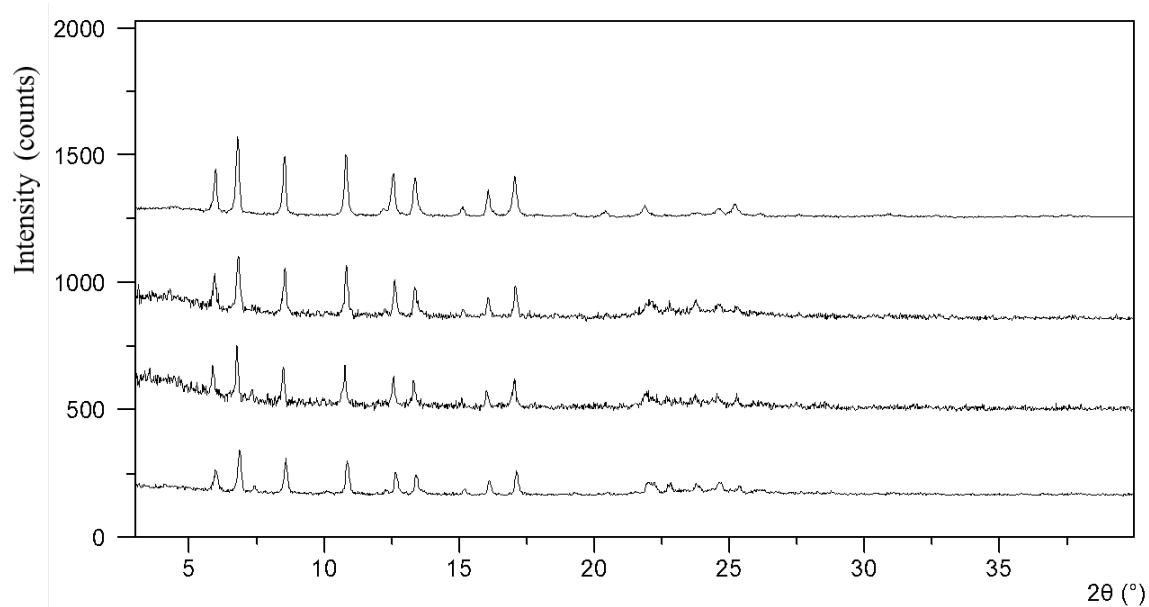


FIG. 6D

D R A W I N G

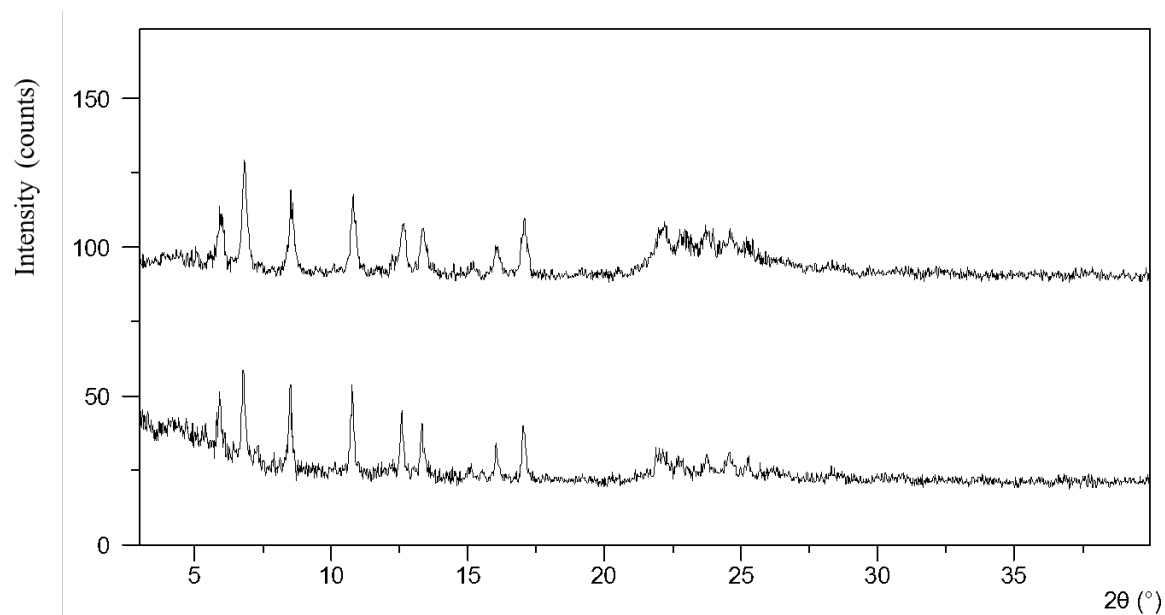


FIG. 6E

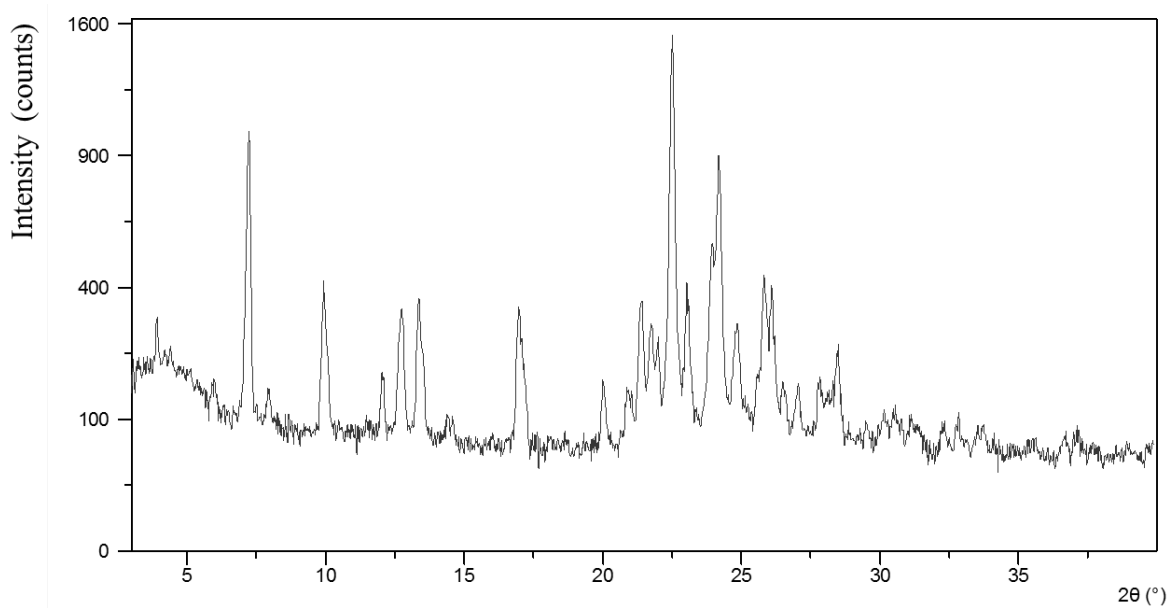


FIG. 7A

D R A W I N G

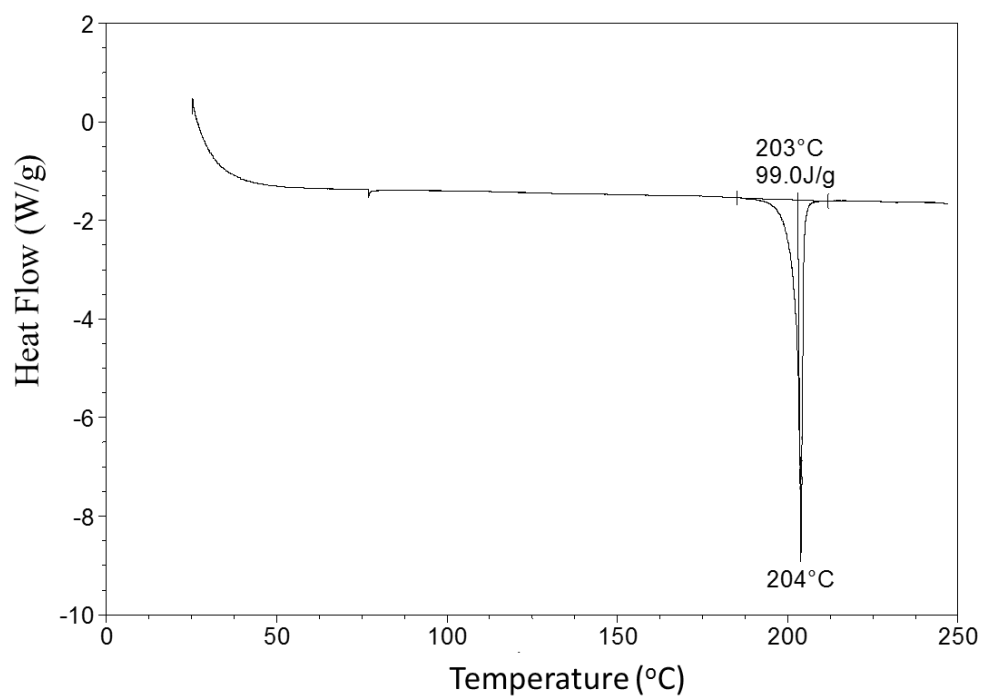


FIG. 7B

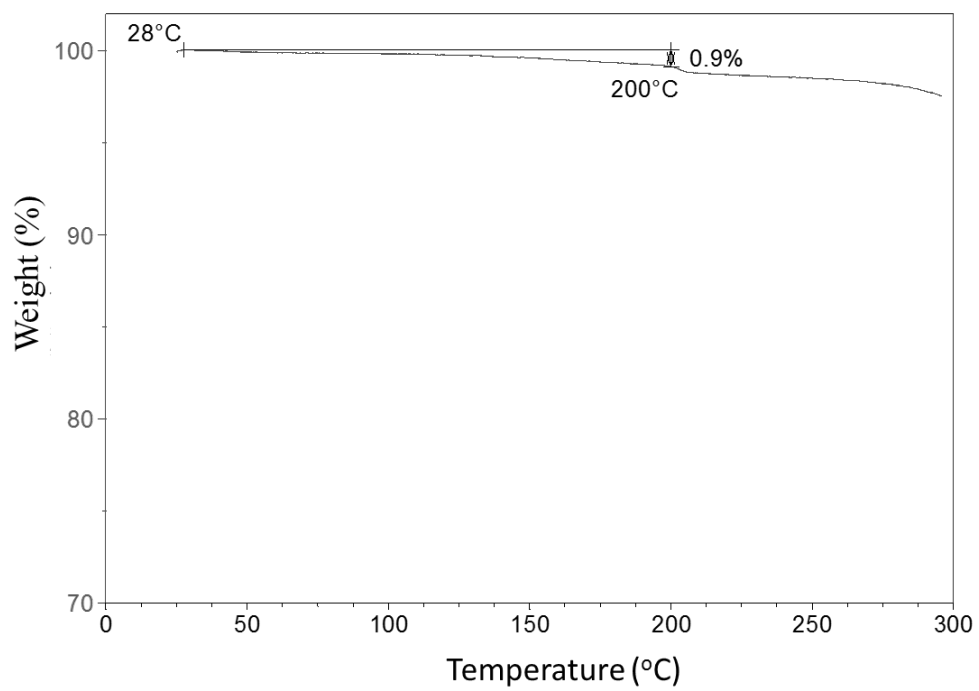


FIG. 7C

DRAWING

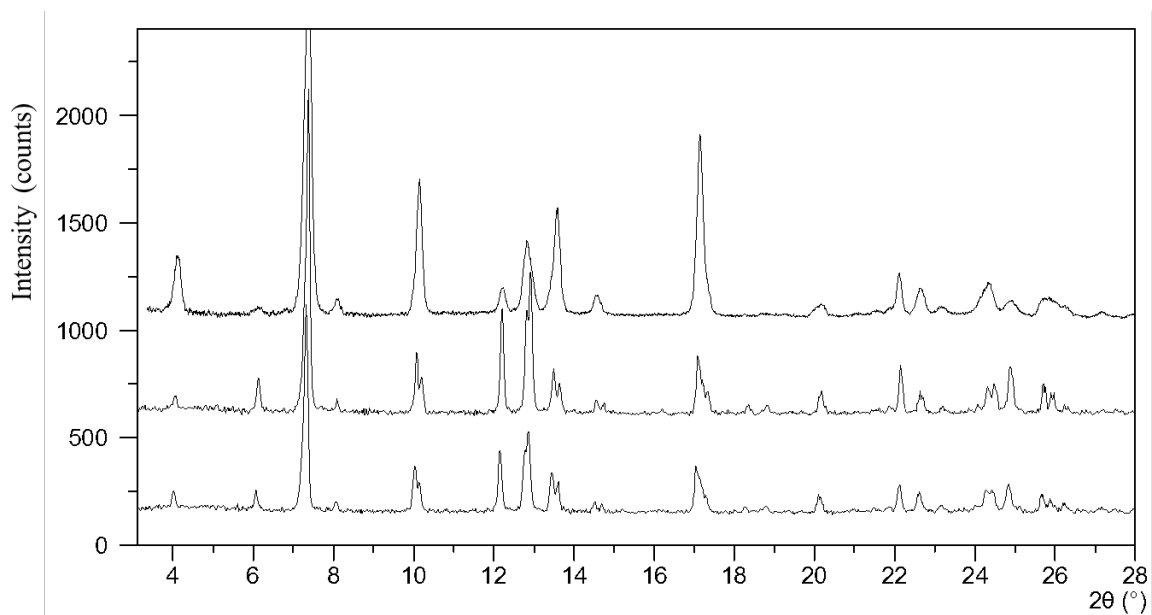


FIG. 7D

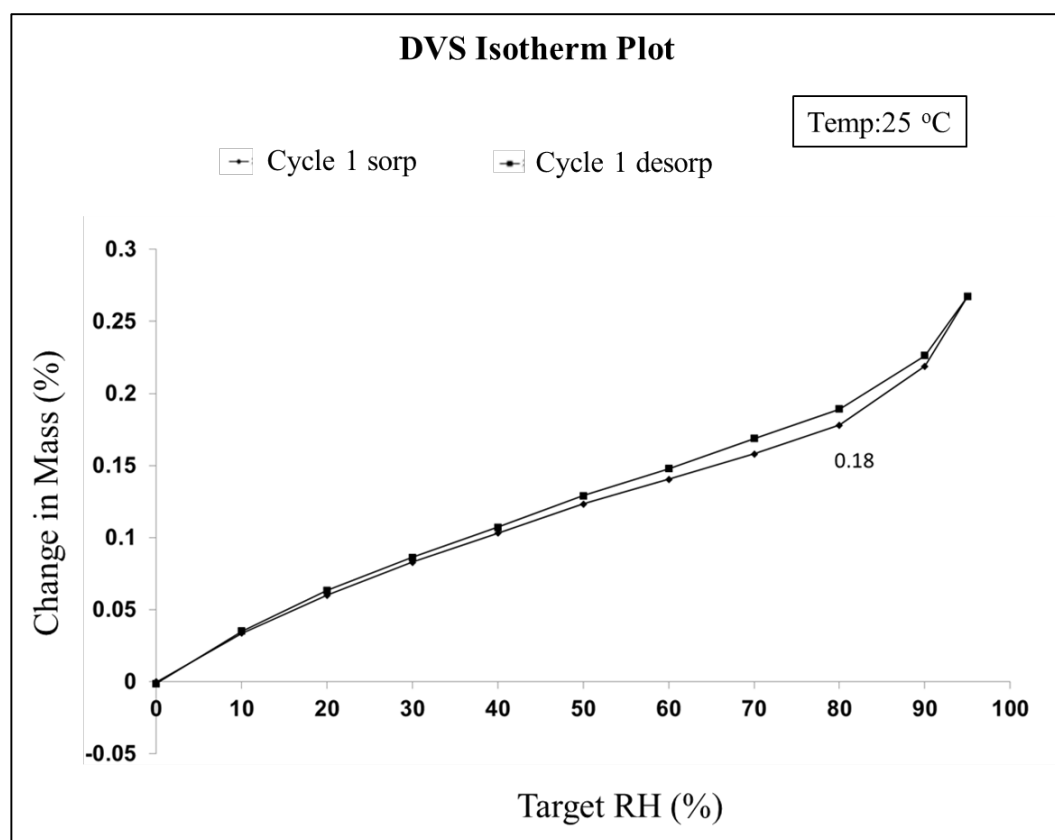


FIG. 7E

DRAWING

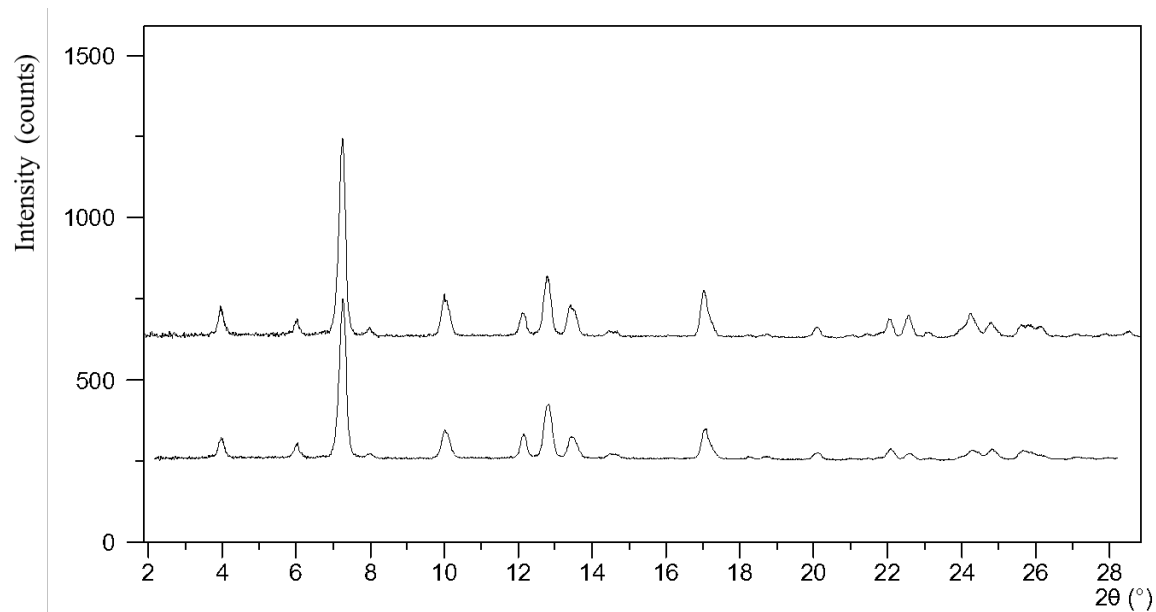


FIG. 7F

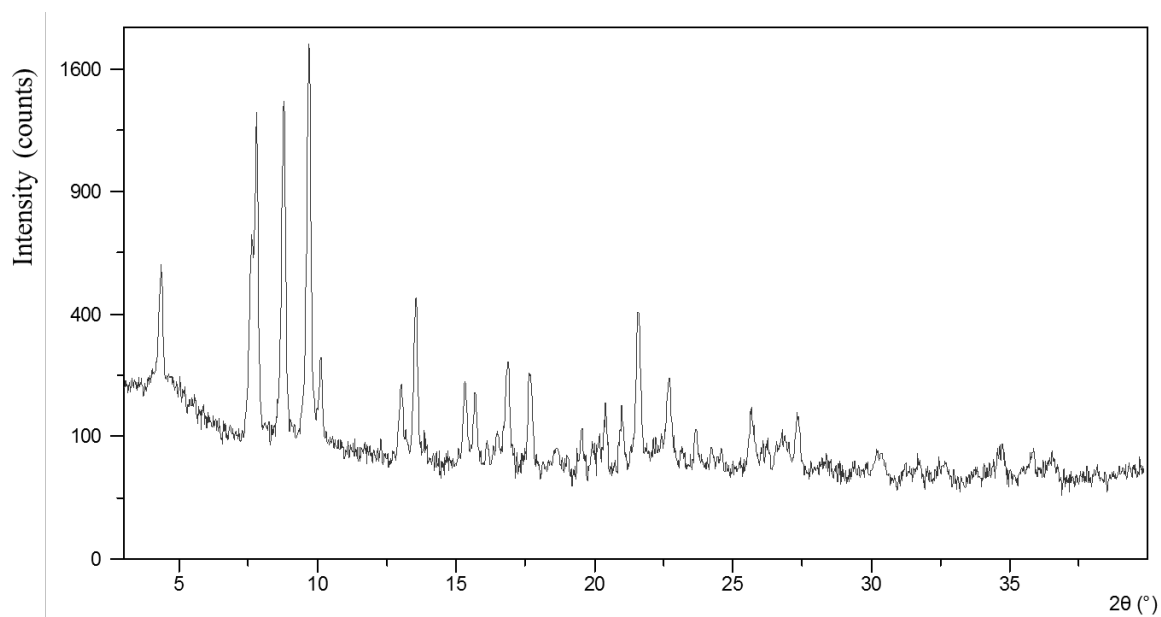


FIG. 8A

DRAWING

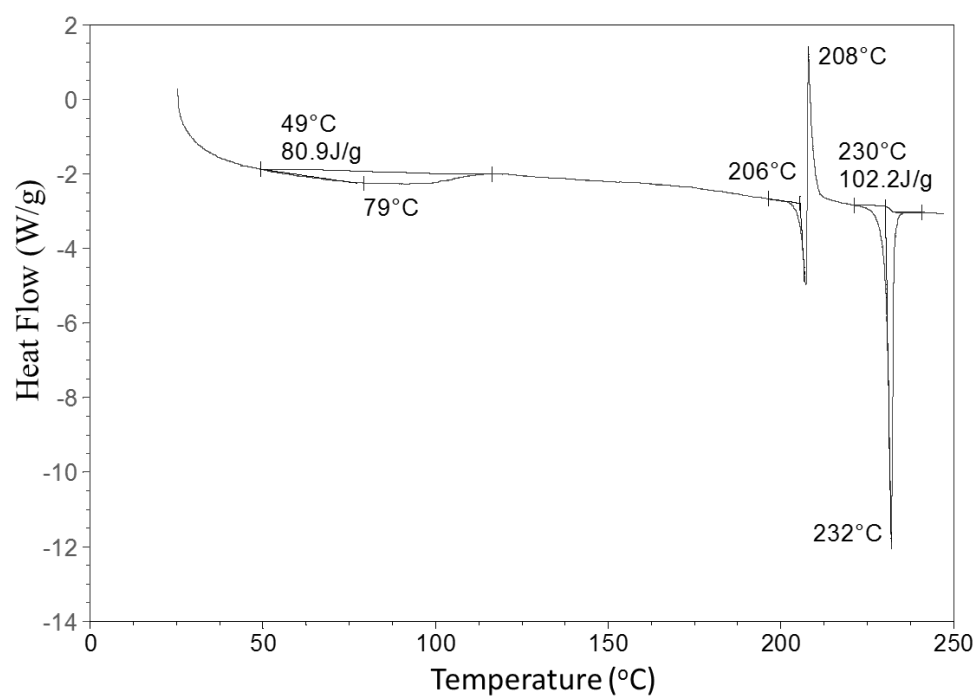


FIG. 8B

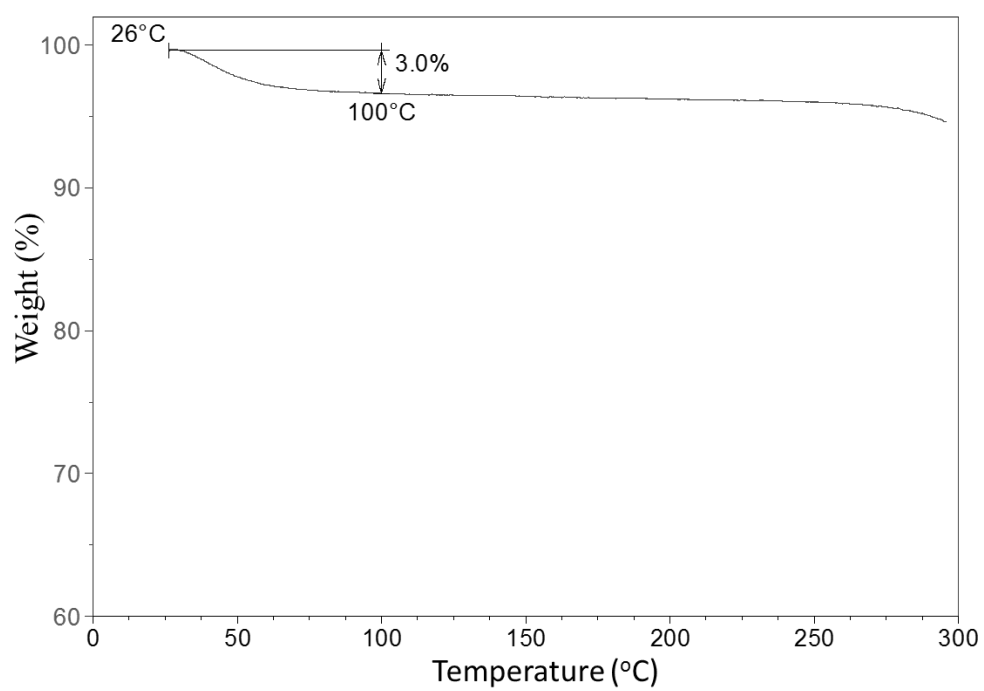


FIG. 8C

DRAWING

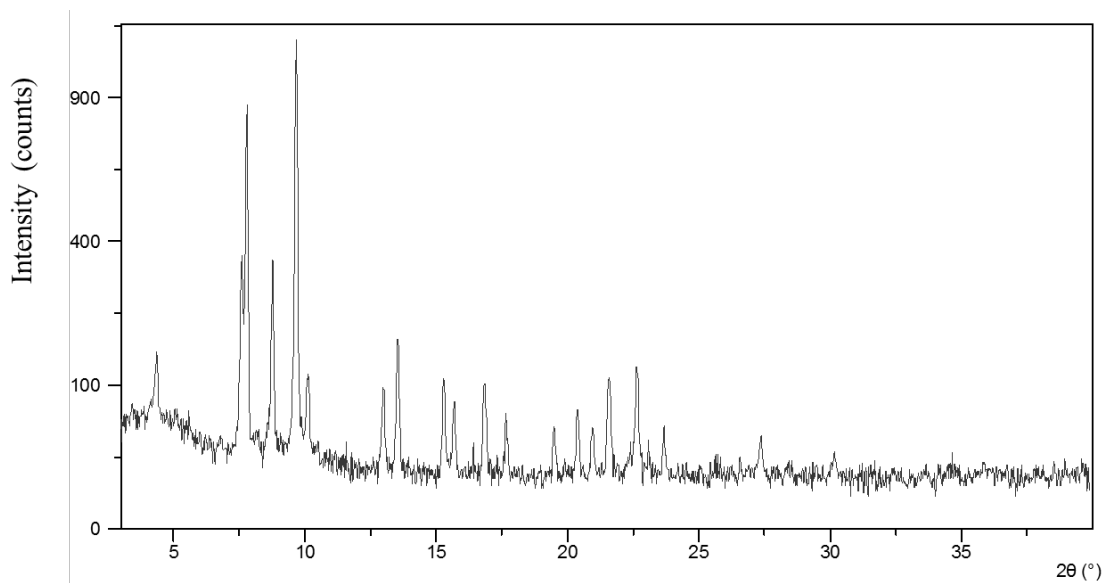


FIG. 8D

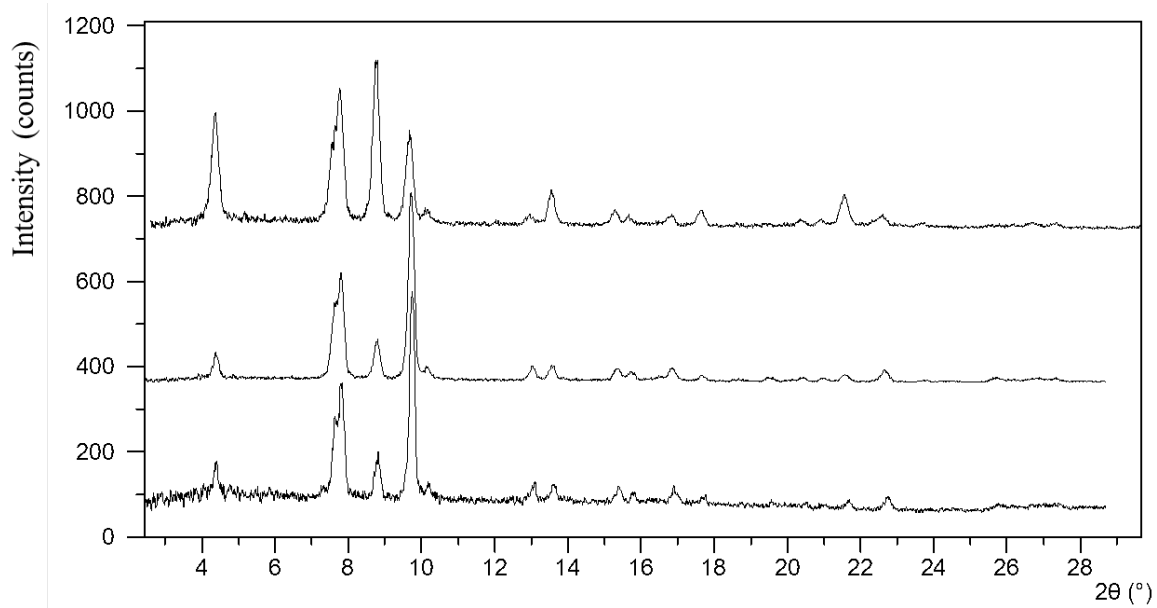


FIG. 8E

DRAWING

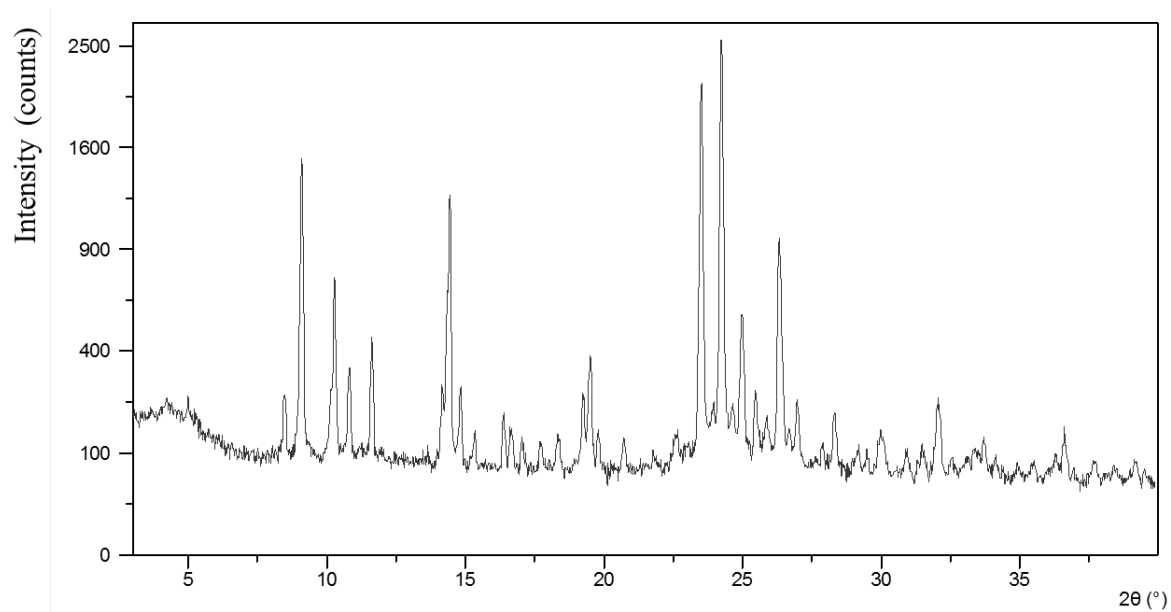


FIG. 9A

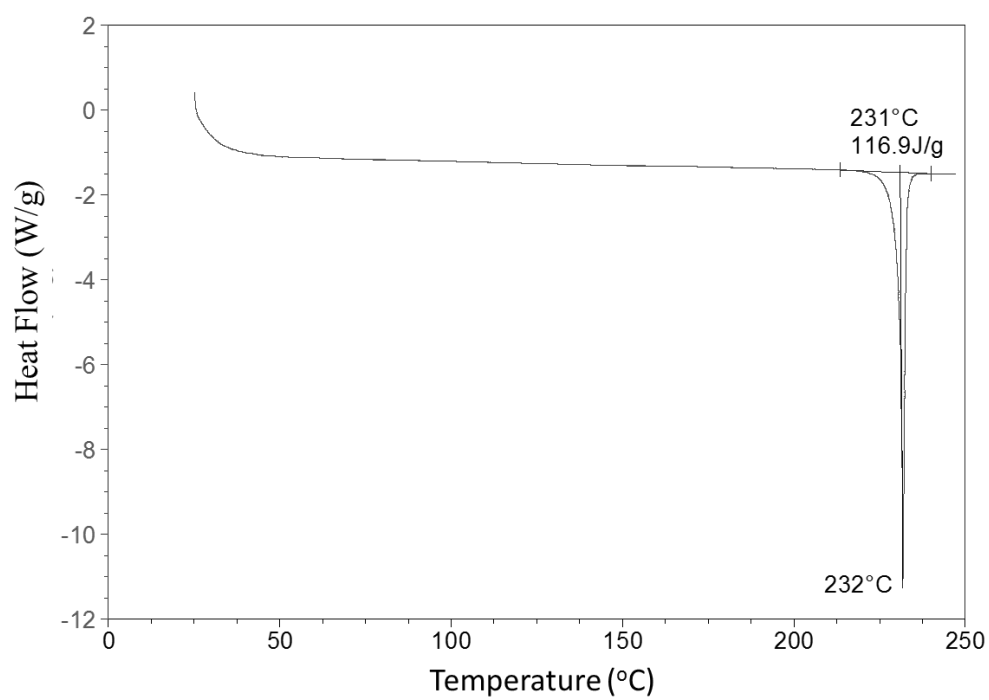


FIG. 9B

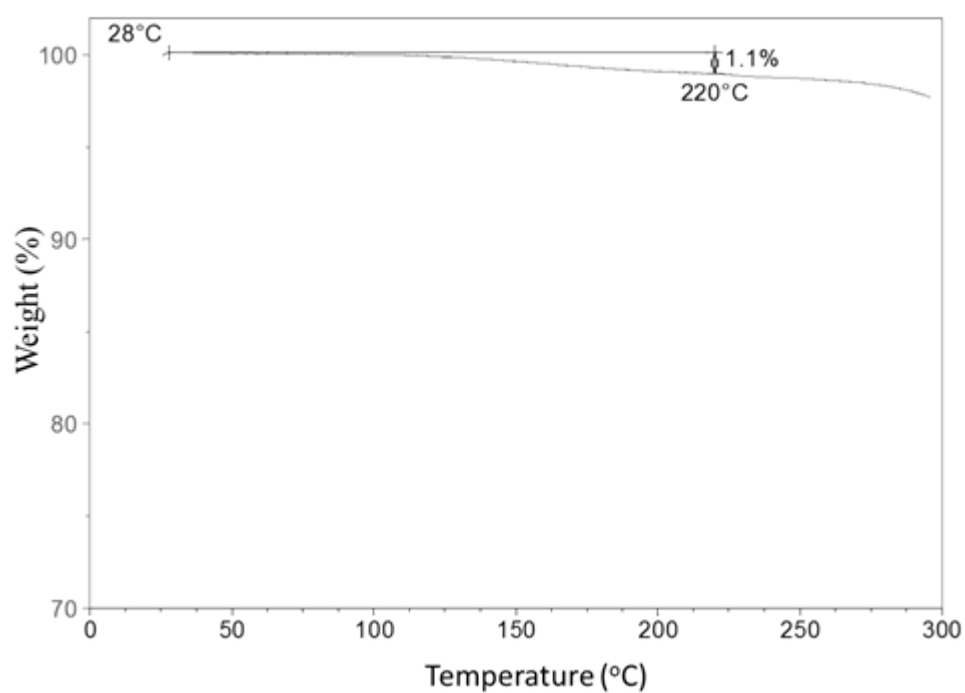


FIG. 9C

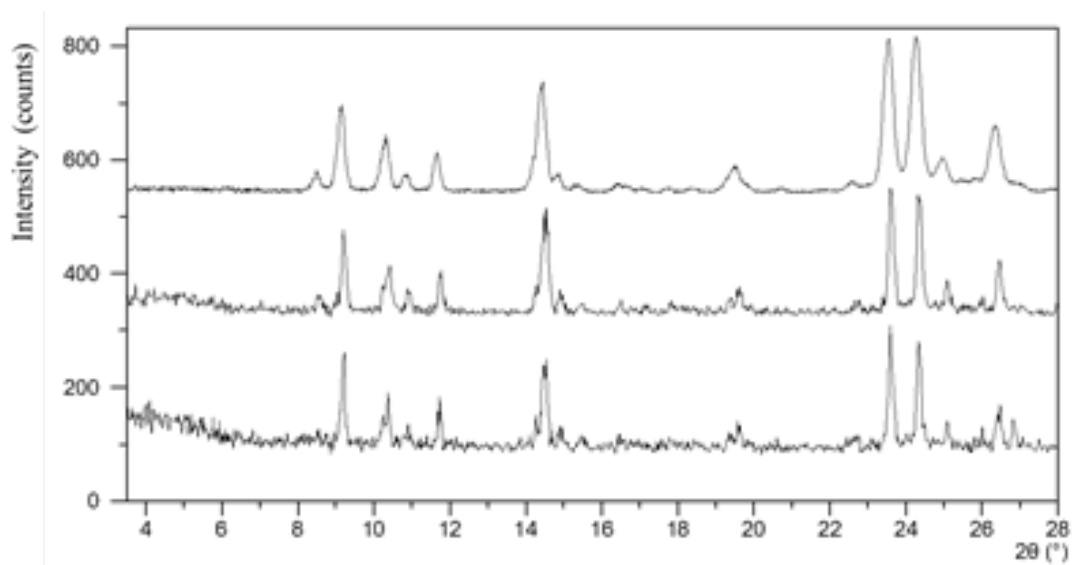


FIG. 9D

D R A W I N G

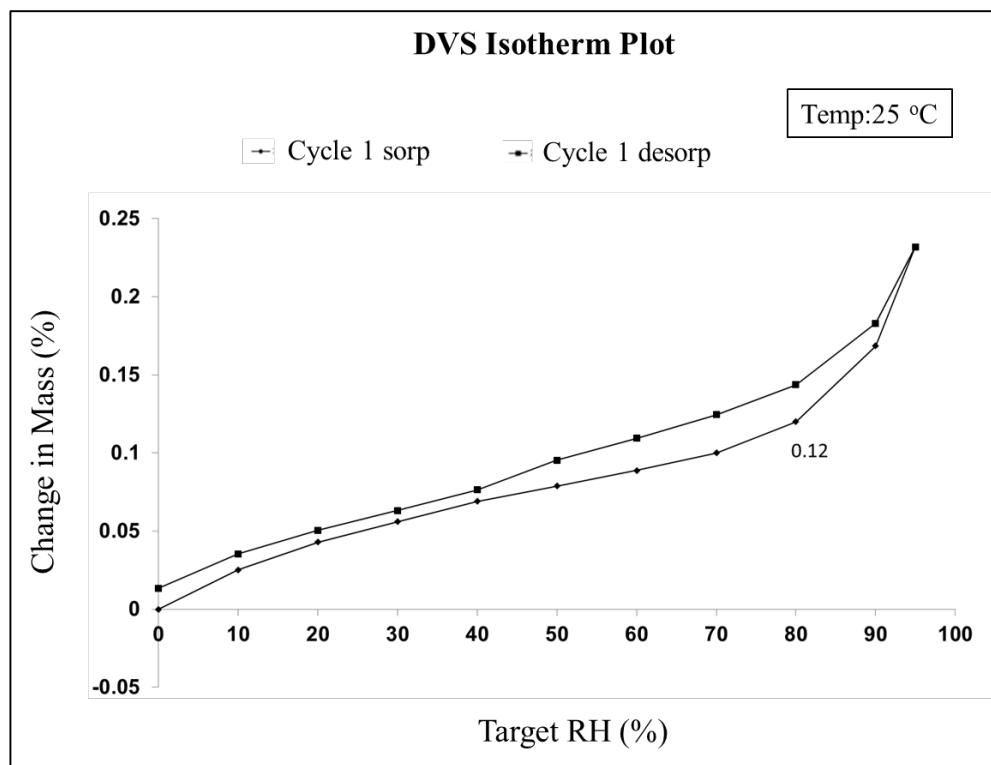


FIG. 9E

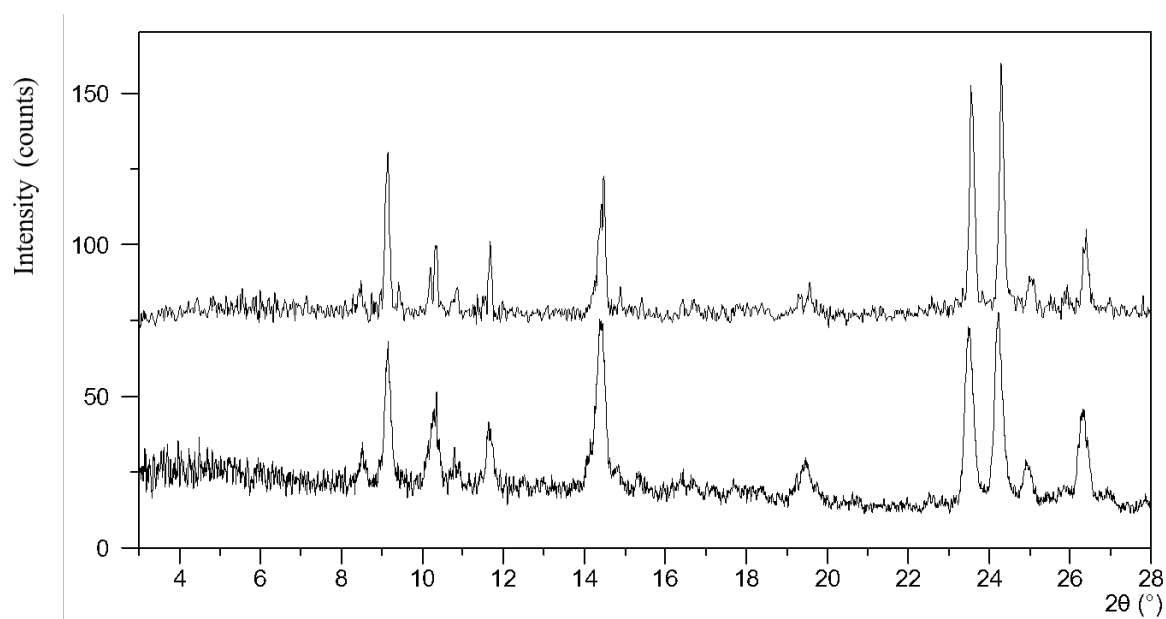


FIG. 9F

D R A W I N G

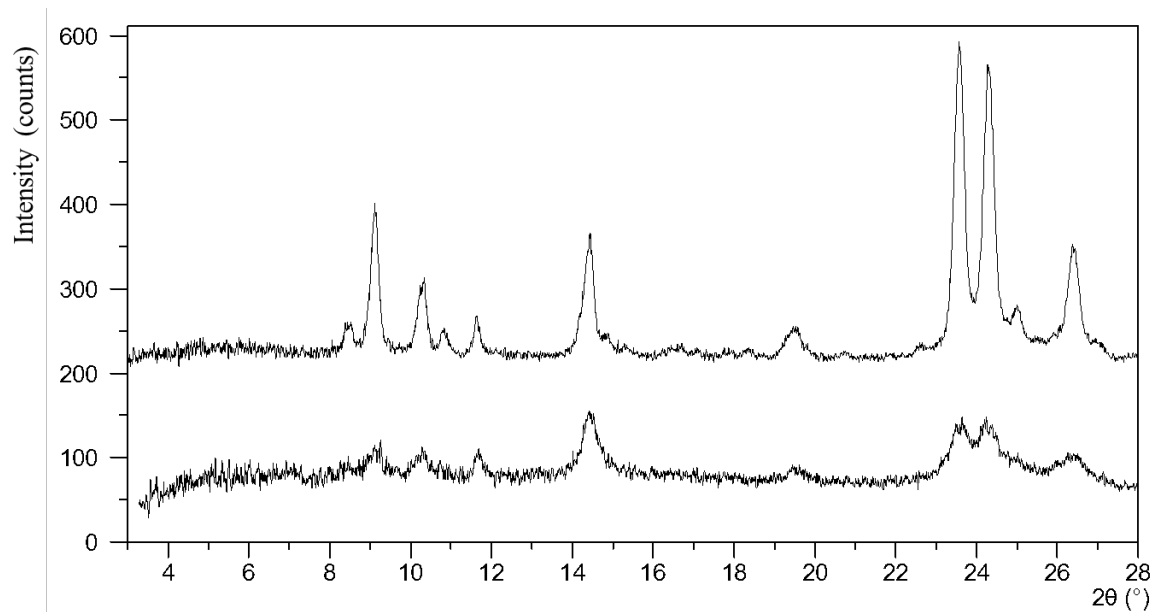


FIG. 9G