



(86) Date de dépôt PCT/PCT Filing Date: 2014/03/21

(87) Date publication PCT/PCT Publication Date: 2014/10/02

(45) Date de délivrance/Issue Date: 2023/07/11

(85) Entrée phase nationale/National Entry: 2015/09/03

(86) N° demande PCT/PCT Application No.: GB 2014/050888

(87) N° publication PCT/PCT Publication No.: 2014/155068

(30) Priorités/Priorities: 2013/03/25 (EP13160777.2);
2013/10/15 (GB1318214.2)

(51) Cl.Int./Int.Cl. *H01M 10/44* (2006.01),
H01M 10/052 (2010.01), *H01M 10/056* (2010.01),
H01M 4/136 (2010.01)

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(54) Titre : PROCEDE DE CHARGE D'UNE BATTERIE AU LITHIUM-SOUFRE

(54) Title: A METHOD OF CHARGING A LITHIUM-SULPHUR CELL

(57) Abrégé/Abstract:

A method for charging a lithium-sulphur cell, said method comprising: monitoring the voltage, V , of a cell during charge as a function of time, t , or capacity, Q , determining, in a voltage region in which the cell transitions between the first stage and second stage of charge, the reference capacity, Q_{ref} , of the cell at which dV/dt or dV/dQ is at a maximum, terminating charge when the capacity of the cell reaches $a \cdot Q_{ref}$, where a is 1.1 to 1.4.

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property
Organization
International Bureau



(43) International Publication Date
2 October 2014 (02.10.2014)

WIPO | PCT

(10) International Publication Number
WO 2014/155068 A1

(51) International Patent Classification:

H01M 4/38 (2006.01) *H01M 10/44* (2006.01)
H01M 10/052 (2010.01)

(21) International Application Number:

PCT/GB2014/050888

(22) International Filing Date:

21 March 2014 (21.03.2014)

(25) Filing Language:

English

(26) Publication Language:

English

(30) Priority Data:

13160777.2 25 March 2013 (25.03.2013) EP
1318214.2 15 October 2013 (15.10.2013) GB

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(81) Designated States (unless otherwise indicated, for every kind of national protection available): AE, AG, AL, AM, AO, AT, AU, AZ, BA, BB, BG, BH, BN, BR, BW, BY, BZ, CA, CH, CL, CN, CO, CR, CU, CZ, DE, DK, DM, DO, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, GT, HN, HR, HU, ID, IL, IN, IR, IS, JP, KE, KG, KN, KP, KR, KZ, LA, LC, LK, LR, LS, LT, LU, LY, MA, MD, ME, MG, MK, MN, MW, MX, MY, MZ, NA, NG, NI, NO, NZ, OM, PA, PE, PG, PH, PL, PT, QA, RO, RS, RU, RW, SA, SC, SD, SE, SG, SK, SL, SM, ST, SV, SY, TH, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, ZA, ZM, ZW.

(84) Designated States (unless otherwise indicated, for every kind of regional protection available): ARIPO (BW, GH, GM, KE, LR, LS, MW, MZ, NA, RW, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, RU, TJ, TM), European (AL, AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HR, HU, IE, IS, IT, LT, LU, LV, MC, MK, MT, NL, NO, PL, PT, RO, RS, SE, SI, SK, SM, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, KM, ML, MR, NE, SN, TD, TG).

Published:

— with international search report (Art. 21(3))

(54) Title: A METHOD OF CHARGING A LITHIUM-SULPHUR CELL

(57) Abstract: A method for charging a lithium-sulphur cell, said method comprising: monitoring the voltage, V , of a cell during charge as a function of time, t , or capacity, Q , determining, in a voltage region in which the cell transitions between the first stage and second stage of charge, the reference capacity, Q_{ref} , of the cell at which dV/dt or dV/dQ is at a maximum, terminating charge when the capacity of the cell reaches $a \cdot Q_{ref}$, where a is 1.1 to 1.4.



WO 2014/155068 A1

A METHOD OF CHARGING A LITHIUM-SULPHUR CELL

[0001] The present invention relates to a method of charging a lithium-sulphur battery. The present invention also relates to a battery management system for charging a lithium-sulphur battery.

BACKGROUND

[0002] A typical lithium-sulphur cell comprises an anode (negative electrode) formed from lithium metal or a lithium metal alloy, and a cathode (positive electrode) formed from elemental sulphur or other electroactive sulphur material. The sulphur or other electroactive sulphur-containing material may be mixed with an electrically conductive material, such as carbon, to improve its electrical conductivity. Typically, the carbon and sulphur are ground and then mixed with a solvent and a binder to form a slurry. The slurry is applied to a current collector and then dried to remove solvent. The resulting structure is calendared to form a composite structure, which is cut into the desired shape to form a cathode. A separator is placed on the cathode and a lithium anode placed on the separator. Electrolyte is then introduced into the assembled cell to wet the cathode and separator.

[0003] Lithium-sulphur cells are secondary cells. When a lithium-sulphur cell is discharged, the sulphur in the cathode is reduced in two-stages. In the first stage, the sulphur (e.g. elemental sulphur) is reduced to polysulphide species, S_n^{2-} ($n \geq 2$). These species are generally soluble in the electrolyte. In the second stage of discharge, the polysulphide species are reduced to lithium sulphide, Li_2S , which, typically, deposits on the surface of the anode.

[0004] When the cell is charged, the two-stage mechanism occurs in reverse, with the lithium sulphide being oxidised to lithium polysulphide and thereafter to lithium and sulphur. This two-stage mechanism can be seen in both the discharging and charging profiles of a lithium-sulphur cell. Accordingly, when a lithium-sulphur cell is charged, its voltage typically passes through an inflexion point as the cell transitions between the first and second stage of charge.

[0005] Lithium-sulphur cells may be (re)charged by applying an external current to the cell. Typically, the cell is charged to a fixed cut-off voltage of, for example, 2.45 V. However, with repeated cycling over an extended period, the capacity of the cell may fade. Indeed, after a certain number of cycles, it may no longer be possible to charge the cell to the fixed cut-off voltage because of the increasing internal resistance of the cell. By repeatedly charging the

cell to the selected cut-off voltage, the cell may eventually be repeatedly over-charged. This can have a detrimental effect on the longevity of the cell, as undesirable chemical reactions may lead to damage to, for example, the cell's electrodes and/or electrolytes.

[0006] In view of the foregoing, it is desirable to avoid over-charging the lithium-sulphur cell. WO 2007/111988 describes a process for determining the point at which a lithium sulphur cell becomes fully charged. Specifically, this reference describes adding an N-O additive, such as lithium nitrate, to the electrolyte of the cell. According to the passage at page 16, lines 29 to 31, of this reference, the additive is effective in providing a charge profile with a sharp increase in voltage at the point of full charge. Accordingly, charging can be terminated once this rapid increase in voltage is observed. This sharp increase in voltage occurs once the second stage of charge is complete, at voltages beyond the inflexion point between the two stages of charge described above.

[0007] The method of WO 2007/111988 relies on the voltage of the cell increasing very sharply as the cell reaches full capacity. Not all lithium-sulphur cells, however, exhibit such a charging profile.

SUMMARY OF THE INVENTION

[0008] According to the present invention, there is provided a method for charging a lithium-sulphur cell, said method comprising:

monitoring the voltage, V , of a cell during charge as a function of time (t) or capacity (Q),

determining in a voltage region in which the cell transitions between the first stage and second stage of charge, the reference capacity, Q_{ref} , of the cell at which dV/dt or dV/dQ is at a maximum,

terminating charge when the capacity of the cell reaches $a.Q_{ref}$, where a is 1.1 to 1.4.

[0009] Preferably, a is 1.2 to 1.3, more preferably 1.25.

Brief description of the Drawings

[0010] Embodiments of the invention will now be described by way of example only with reference to the accompanying drawings wherein:

[0011] Figure 1: is a charging profile of a lithium-sulphur cell, showing the inflexion point between the two stages of charge.

[0012] Figure 2: is a graph of dV/dt (V/min) over time.

[0013] Figure 3: is a graph of $dV^2/d^2T(V^2/min)$ over time

Detailed Description

[0014] When a lithium-sulphur cell is discharged, the sulphur material present in the cathode is reduced in two-stages. In particular, the sulphur is first reduced to polysulphide species, S_n^{2-} ($n \geq 2$), in the first stage of discharge. These polysulphide species are generally soluble in the electrolyte. In the second stage of discharge, the polysulphide species are reduced to lithium sulphide, Li_2S , which deposits on the surface of the lithium anode.

[0015] When the cell is charged, the two-stage mechanism occurs in reverse, with lithium sulphide being oxidised to lithium polysulphide and thereafter to lithium and sulphur. This two-stage mechanism can be seen in both the discharging and charging profiles of lithium-sulphur cells. Thus, when a lithium-sulphur cell is charged, its voltage passes through an inflexion point with time, as the cell transitions between the first and second stage of charge. Figure 1 is a charging profile of a lithium-sulphur cell, showing the inflexion point between the two stages of charge.

[0016] In the present invention, the reference capacity, Q_{ref} , at the inflexion point is determined by determining the capacity at which dV/dt or dV/dQ is at a maximum. The inflexion point occurs in the voltage region where the cell transitions between the first stage and second stage of charge. Thus, the inflexion point does not occur beyond the second stage of charge as the cell approaches full charge. Instead, the inflexion point is advantageously a reference point that occurs between the first and second stages of charge from which the termination voltage of a cell at a particular charging cycle can be determined. Typically, the inflexion point occurs when a significant proportion of the cathodic sulphur material (e.g. elemental sulphur) is still dissolved in the electrolyte (e.g. as polysulphide). Preferably, the inflexion point occurs when at least 80%, more preferably at least 90%, yet more preferably at least 95% of the cathodic sulphur material is dissolved in the electrolyte (e.g. as polysulphide). The percentage of cathodic sulphur material dissolved in solution can be determined by known methods, for example, from the amount of residual solid sulphur in a cell as a percentage of the initial amount of sulphur material introduced as the cathodic material.

[0017] In a preferred embodiment, the point at which dV/dt or dV/dQ is a maximum is determined at a voltage less than 2.4 V. Typically, the cell transitions between the first stage and second stage of discharge between 1.5 to 2.4V, preferably 1.8 to 2.38 V, more preferably 1.9 to 2.35 V, for example, 2.0 to 2.35 V. The value of dV/dt or dV/dQ , therefore,

is advantageously calculated within such voltage ranges. In one embodiment, the value of dV/dt or dV/dQ is calculated at selected points within such voltage ranges. For the avoidance of doubt, the capacity of the cell, Q_{ref} , at which dV/dt or dV/dQ is a maximum can be determined by methods that are well-known in the art. For example, the capacity, Q , of the cell may be determined by multiplying current by time in the case where the cell is charged using a constant current. Even if dV/dQ is determined, therefore, the value of Q_{ref} can be readily calculated.

[0018] In one embodiment, the reference capacity, Q_{ref} , is determined by calculating dV/dt or dV/dQ of the charging cycle as the voltage of the cell increases, for example, to a value of 2.4 or lower, preferably 2.38 V or lower, more preferably 2.35 V or lower. In one embodiment, the reference capacity, Q_{ref} , is determined by calculating dV/dt or dV/dQ of the charging cycle as the voltage of the cell increases, for example, from a value of 1.5 V or higher, preferably 1.8 V or higher, more preferably 1.9 V or higher, for example, 2.0 V or higher. In a preferred embodiment, the reference capacity, Q_{ref} , is determined by calculating dV/dt or dV/dQ , for example, at least in a selected voltage range of 1.5 to 2.4V, preferably 1.8 to 2.38 V, more preferably 1.9 to 2.35, even more preferably from 2.0 to 2.35 V, yet more preferably 2.1 to 2.35, for example, 2.2 to 2.35. In one embodiment, the reference capacity, Q_{ref} , is determined by calculating dV/dt or dV/dQ , for example, in a selected voltage range of 2.25 to 2.35. The values of dV/dt or dV/dQ may be determined at selected points within a range and the maximum value of dV/dt or dV/dQ may be determined or calculated from these selected points.

[0019] Because the reference capacity, Q_{ref} , will vary depending on the specific charge cycle of the cell, the cut-off voltage of the cell will vary depending on the current or contemporaneous characteristics (e.g. voltage vs time charging profile) of the cell. For example, as the cell ages, the characteristics (e.g. voltage vs time charging profile) of the cell may change and, as a result, so will the value of Q_{ref} and the cut-off or termination point as determined by the method of the present invention. The risk of over-charging a cell using the method of the present invention, therefore, is reduced, as the cut-off point can be determined depending on the current or contemporaneous characteristics of the cell during each individual charge cycle. Advantageously, this reduces the risk of capacity fade. In contrast, in conventional methods of charging a lithium sulphur cell, the cell is charged to a pre-determined voltage or capacity irrespective of the charging characteristics of the cell and the extent to which the cell has experienced capacity fade. The risk of overcharging the cell, therefore, is significant, particularly as the cell ages.

[0020] In one embodiment of the present invention, it may be desirable to calculate dV/dt or dV/dQ at a plurality of points, for example, in a selected voltage range. The values of d^2V/dt^2 or d^2V/dQ^2 may also be determined to ascertain when d^2V/dt^2 or d^2V/dQ^2 is zero in order to confirm that dV/dt or dV/dQ is indeed at its maximum. Preferably, the charge is terminated when dV/dt or dV/dQ is at its maximum and d^2V/dt^2 or d^2V/dQ^2 is zero. In one embodiment, the charge is terminated when dV/dt is at its maximum and d^2V/dt^2 is zero. In another embodiment, the charge is terminated when dV/dQ is at its maximum and d^2V/dQ^2 is zero. For the avoidance of doubt, the point at which dV/dt or dV/dQ is at its maximum and d^2V/dt^2 or d^2V/dQ^2 is zero occurs as the cell transitions between the first and second stage of charge as discussed above.

[0021] In another aspect of the present invention, there is provided a battery management system for carrying out the method as described above. In yet another aspect, there is provided a battery management system for a lithium-sulphur battery, said system comprising:

means for monitoring the voltage, V , of a cell during charge as a function of time, t , or capacity, Q ,

means for determining in the voltage region where the cell transitions between the first stage and second stage of charge, a reference capacity, Q_{ref} , of the cell at which dV/dt is at a maximum, and,

means for terminating charge when the capacity of the cell reaches $a.Q_f$, where a is 1.1 to 1.4.

[0022] The system may additionally include means for coupling the system to a lithium-sulphur battery. Preferably, the system comprises a lithium sulphur battery.

[0023] In a preferred embodiment, the lithium-sulphur cell is charged by supplying electric energy at constant current. The current may be supplied so as to charge the cell in a time ranging from 30 minutes to 12 hours, preferably 8 to 10 hours. The current may be supplied at a current density ranging from 0.1 to 3 mA/cm², preferably 0.1 to 0.3 mA/cm². As an alternative to charging at a constant current, it may also be possible to charge the lithium-sulphur cell to a constant voltage until the relevant capacity is reached. Suitable voltages range from 2.4-3.0V.

[0024] The electrochemical cell may be any suitable lithium-sulphur cell. The cell typically includes an anode, a cathode, an electrolyte and, advantageously, a porous separator positioned between the anode and cathode. The anode may be formed of lithium metal or a

lithium metal alloy. Preferably, the anode is a metal foil electrode, such as a lithium foil electrode. The lithium foil may be formed of lithium metal or lithium metal alloy.

[0025] The cathode of the electrochemical cell includes a mixture of electroactive sulphur material and electroconductive material. This mixture forms an electroactive layer, which may be placed in contact with a current collector.

[0026] The mixture of electroactive sulphur material and electroconductive material may be applied to the current collector in the form of a slurry in a solvent (e.g. water or an organic solvent). The solvent may then be removed and the resulting structure calendared to form a composite structure, which may be cut into the desired shape to form a cathode. A separator may be placed on the cathode and a lithium anode placed on the separator. Electrolyte may then be introduced into the assembled cell to wet the cathode and separator.

[0027] The electroactive sulphur material may comprise elemental sulphur, sulphur-based organic compounds, sulphur-based inorganic compounds and sulphur-containing polymers. Preferably, elemental sulphur is used.

[0028] The solid electroconductive material may be any suitable conductive material. Preferably, this solid electroconductive material may be formed of carbon. Examples include carbon black, carbon fibre and carbon nanotubes. Other suitable materials include metal (e.g. flakes, filings and powders) and conductive polymers. Preferably, carbon black is employed.

[0029] The weight ratio of electroactive sulphur material (e.g. elemental sulphur) to electroconductive material (e.g. carbon) may be 1 to 30:1; preferably 2 to 8:1, more preferably 5 to 7:1.

[0030] The mixture of electroactive sulphur material and electroconductive material may be a particulate mixture. The mixture may have an average particle size of 50 nm to 20 microns, preferably 100 nm to 5 microns.

[0031] The mixture of electroactive sulphur material and electroconductive material (i.e. the electroactive layer) may optionally include a binder. Suitable binders may be formed from at least one of, for example, polyethylene oxide, polytetrafluoroethylene, polyvinylidene fluoride, ethylene-propylene-diene rubber, methacrylate (e.g. UV-curable methacrylate), and divinyl esters (e.g. heat curable divinyl esters).

[0032] As discussed above, the cathode of the electrochemical cell may further comprise a current collector in contact with the mixture of electroactive sulphur material and solid electroconductive material. For example, the mixture of electroactive sulphur material and

solid electroconductive material is deposited on the current collector. A separator is also disposed between the anode and the cathode of the electrochemical cell. For example, the separator may be in contact with the mixture of electroactive sulphur material and solid electroconductive material, which, in turn, is in contact with the current collector.

[0033] Suitable current collectors include metal substrates, such as foil, sheet or mesh formed of a metal or metal alloy. In a preferred embodiment, the current collector is aluminium foil.

[0034] The separator may be any suitable porous substrate that allows ions to move between the electrodes of the cell. The porosity of the substrate should be at least 30%, preferably at least 50%, for example, above 60%. Suitable separators include a mesh formed of a polymeric material. Suitable polymers include polypropylene, nylon and polyethylene. Non-woven polypropylene is particularly preferred. It is possible for a multi-layered separator to be employed.

[0035] Preferably, the electrolyte comprises at least one lithium salt and at least one organic solvent. Suitable lithium salts include at least one of lithium hexafluorophosphate (LiPF_6), lithium hexafluoroarsenate (LiAsF_6), lithium perchlorate (LiClO_4), lithium trifluoromethanesulfonimide ($\text{LiN}(\text{CF}_3\text{SO}_2)_2$), lithium borofluoride and lithium trifluoromethanesulphonate ($\text{CF}_3\text{SO}_3\text{Li}$). Preferably the lithium salt is lithium trifluoromethanesulphonate.

[0036] Suitable organic solvents are tetrahydrofuran, 2-methyltetrahydrofuran, dimethylcarbonate, diethylcarbonate, ethylmethylcarbonate, methylpropylcarbonate, methylpropylpropionate, ethylpropylpropionate, methyl acetate, dimethoxyethane, 1, 3-dioxolane, diglyme (2-methoxyethyl ether), tetraglyme, ethylene carbonate, propylene carbonate, γ -butyrolactone, dioxolane, hexamethyl phosphoamide, pyridine, dimethyl sulfoxide, tributyl phosphate, trimethyl phosphate, N, N, N, N-tetraethyl sulfamide, and sulfone and their mixtures. Preferably, the organic solvent is a sulfone or a mixture of sulfones. Examples of sulfones are dimethyl sulfone and sulfolane. Sulfolane may be employed as the sole solvent or in combination, for example, with other sulfones.

[0037] The organic solvent used in the electrolyte should be capable of dissolving the polysulphide species, for example, of the formula S_n^{2-} , where $n = 2$ to 12, that are formed when the electroactive sulphur material is reduced during discharge of the cell.

[0038] The concentration of lithium salt in the electrolyte is preferably 0.1 to 5M, more preferably 0.5 to 3M, for example, 1M. The lithium salt is preferably present at a

concentration that is at least 70%, preferably at least 80%, more preferably at least 90%, for example, 95 to 99% of saturation.

[0039] In one embodiment, the electrolyte comprises lithium trifluoromethanesulphonate and sulfolane.

[0040] The weight ratio of electrolyte to the total amount of electroactive sulphur material and electroconductive material is 1 – 15 : 1; preferably 2 – 9 : 1, more preferably 6 - 8 : 1.

Examples

In this example, the voltage, V , of a cell during charge as a function of time (t) was monitored. The voltage was plotted as a function of time in Figure 1. As can be seen from Figure 1, the cell transitions between a first stage of charge to a second stage of charge in the region between about 2.2 V and 2.4 V. In this region, a reference capacity, Q_{ref} , of the cell is determined by calculating dV/dt at a range of points. The reference capacity, Q_{ref} , is determined at the point at which dV/dt is at a maximum, and, d^2V/dt^2 is zero (see Figures 2 and 3). This is the inflexion point between the first and second stages of discharge. In this Example, charging of the cell was terminated when the capacity reached $1.25Q_{ref}$. This process may be repeated for each charge cycle of the cell. In this way, the termination voltage will be determined according to the specific characteristics (e.g. charging profile) of the cell at the charge cycle in question. This reduces the risk of over-charging, as the termination voltage will vary depending on the extent of capacity fade.

Claims

1. A method for charging a lithium-sulphur cell, said method comprising:
 monitoring a voltage, V , of the cell during charge as a function of time, t , or capacity, Q ,
 determining, in a voltage region in which the cell transitions between a first stage and a second stage of charge, a reference capacity, Q_{ref} , of the cell at which dV/dt or dV/dQ is at a maximum,
 terminating charge when the capacity of the cell reaches $a.Q_{ref}$, where a is 1.1 to 1.4.
2. The method as claimed in claim 1, wherein a is 1.2 to 1.3.
3. The method as claimed in claim 2, wherein a is 1.25.
4. The method as claimed in any one of claims 1 to 3, wherein the voltage region at which the cell transitions between the first stage of charge and the second stage of charge is 1.5 to 2.4 V.
5. The method as claimed in claim 4, wherein the voltage region at which the cell transitions between the first stage of charge and the second stage of charge is 1.9 to 2.35 V.
6. The method as claimed in any one of claims 1 to 5, wherein the reference capacity, Q_{ref} , is determined by calculating, in the voltage region at which the cell transitions between the first stage and second stage of charge, the point at which dV/dt or dV/dQ is at a maximum and d^2V/dt^2 or d^2V/dQ^2 is zero.
7. The method as claimed in any one of claims 1 to 6, wherein the reference capacity, Q_{ref} , is determined by calculating dV/dt or dV/dQ in a selected voltage range of 1.5 to 2.4V.

8. The method as claimed in claim 7, wherein the reference capacity, Q_{ref} , is determined by calculating dV/dt or dV/dQ in a selected voltage range of 1.9 to 2.35 V.
9. The method as claimed in any one of claims 1 to 8, which is repeated for at least two charge cycles of the lithium-sulphur cell.
10. The method as claimed in any one of claims 1 to 8, which is repeated for each charge cycle of the lithium-sulphur cell.
11. A battery management system for a lithium-sulphur cell, comprising:
 - means for monitoring a voltage, V , of the cell during charge as a function of time, t , or capacity, Q ,
 - means for determining, in a voltage region where the cell transitions between a first stage and a second stage of charge, a reference capacity, Q_{ref} , of the cell at which dV/dt or dV/dQ is at a maximum,
 - means for terminating charge when the capacity of the cell reaches $a.Q_f$, where a is 1.1 to 1.4 and
 - a lithium sulphur battery.
12. The system as claimed in claim 11, wherein the means for determining the reference capacity, Q_{ref} , are means for determining the reference capacity, Q_{ref} at which dV/dt or dV/dQ is at a maximum and d^2V/dt^2 or d^2V/dQ^2 is zero.
13. The system as claimed in any one of claims 11 to 12, which additionally includes means for coupling the system to the lithium-sulphur battery.

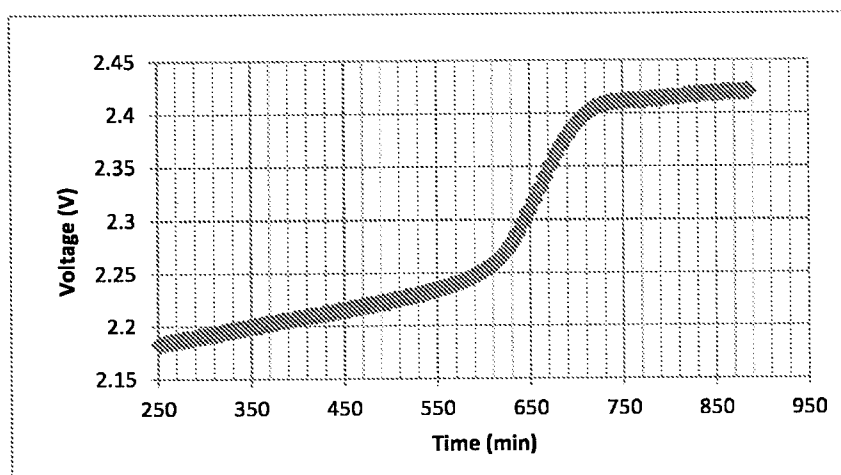


Figure 1

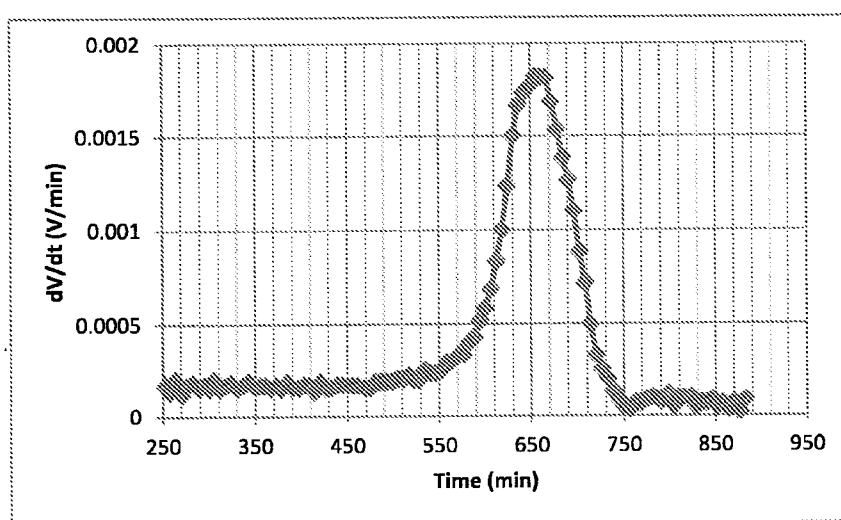


Figure 2

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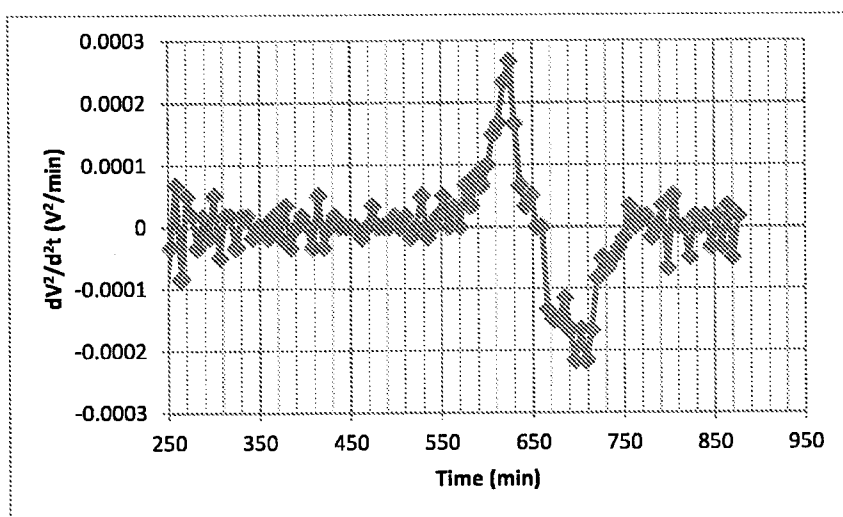


Figure 3