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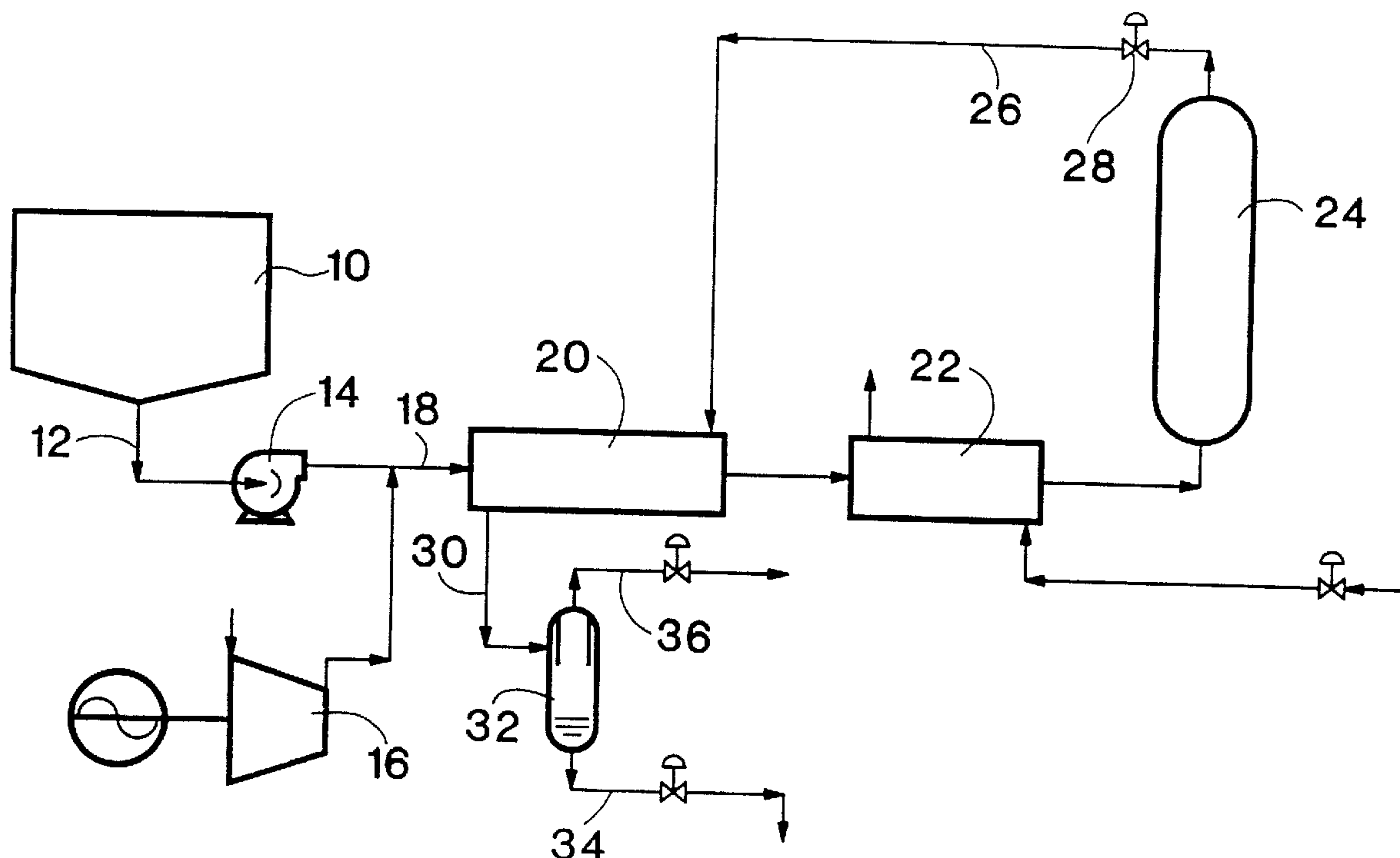
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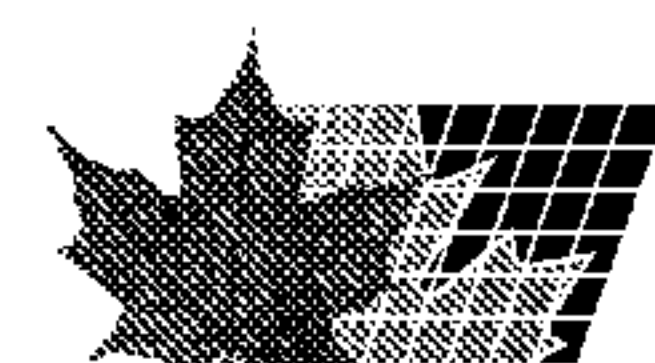
(54) Titre : PROCEDE D'OXYDATION EN MILIEU LIQUIDE DU SULFURE CAUSTIQUE A BASSE TEMPERATURE

(54) Title: LOW TEMPERATURE CAUSTIC SULFIDE WET OXIDATION PROCESS



(57) Abrégé/Abstract:

A process for treatment of caustic sulfide liquor by wet oxidation in a ferrous-based alloy system at temperatures less than about 175°C is described. The liquor is first analyzed for alkalinity consuming species and for nonsulfidic alkalinity. If excess alkalinity consuming species is present compared to nonsulfidic alkalinity, then additional nonsulfidic alkalinity is added to the raw liquor such that excess alkalinity is present during wet oxidation treatment, thus preventing excessive corrosion to the ferrous-base alloy system.



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LOW TEMPERATURE CAUSTIC SULFIDE WET OXIDATION PROCESS

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ABSTRACT

A process for treatment of caustic sulfide liquor by wet
5 oxidation in a ferrous-based alloy system at temperatures less
than about 175°C is described. The liquor is first analyzed
for alkalinity consuming species and for nonsulfidic
alkalinity. If excess alkalinity consuming species is present
compared to nonsulfidic alkalinity, then additional
10 nonsulfidic alkalinity is added to the raw liquor such that
excess alkalinity is present during wet oxidation treatment,
thus preventing excessive corrosion to the ferrous-base alloy
system.

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LOW TEMPERATURE CAUSTIC SULFIDE WET OXIDATION PROCESS

FIELD OF THE INVENTION

This invention relates to a process for treatment of caustic sulfide liquor by wet oxidation, and more particularly to treatment of the liquor in a ferrous-based alloy wet oxidation system.

DESCRIPTION OF RELATED ART

Caustic sulfide liquors are generated in a variety of processes from a number of industrial sources. These caustic liquors are commonly used to remove acidic components from various gas streams by scrubbing or from liquid streams by extraction. The gas streams are generated in the petrochemical industry, petroleum refining, pulp and paper manufacture and various chemical manufacturing processes. The acidic components include hydrogen sulfide, H_2S , mercaptans, RSH , and possibly organic acids. Analysis of caustic sulfide liquors for acidic components, as well as metal carbonates/bicarbonates and caustic soda, is described by J.W. McCoy in Chemical Analysis of Industrial Water, Chemical Publishing Co., New York, 1969. See particularly pages 125-153.

The contaminated caustic liquors represent a formidable disposal problem due to their caustic content as well as the captured acidic components therein. Neutralization of the caustic liquor by acid addition can result in release of the acidic components. Therefore it is essential to convert the captured acidic components to a form suitable for release to

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the environment. Further, there may be additional components present in the caustic liquor which add to the Chemical Oxygen Demand (COD) of the liquor. These components include various carbonaceous materials including phenolics, organic acids and
5 oils.

Wet oxidation is the preferred method of treatment for caustic sulfide liquors since the products of oxidation are inorganic sulfate, carbon dioxide and water. Also, the oxidation is carried out within a closed system which prevents
10 transfer of pollutants to the atmosphere. The highly alkaline nature of these caustic liquors requires special materials of construction for wet oxidation systems employed in their treatment. The nickel-based alloys, such as Inconel 600, are well suited to withstand the elevated temperatures and
15 pressures employed in the wet oxidation process for caustic liquor treatment. The ferrous-based alloys, such as carbon steel, stainless steel, or superstainless steel may be considered for wet oxidation treatment of caustic wastewaters. However, under wet oxidation conditions, corrosion is a
20 serious problem for these alloys. Ferrous-based alloys shall be defined as those containing a majority of iron, at least 50%.

In U.S. Patent No. 3,761,409 McCoy et al. disclose a continuous process for the air oxidation of sulfidic,
25 ammoniacal sour water where feed water is adjusted to a pH between about 6 to 13 and the oxidation occurs at 250°F to 520°F at 75 to 800 psig with up to 500% excess oxygen based on the stoichiometric conversion of sulfide to sulfate.

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Chowdhury in U.S. Patent No. 4,350,599 discloses wet oxidation of caustic liquor where carbon dioxide generated by the oxidation is used to reduce the pH of the caustic feed liquor to below 11. Maintaining the feed below pH 11.0 but
5 above 7.0 prevents corrosion of the less expensive stainless steel wet oxidation system.

Applicants in U.S. Patent No. 5,082,571 have devised a process which relates the species present in the caustic liquor to the amount of caustic required to maintain an excess
10 of alkalinity in the liquor during wet oxidation treatment. This process allows a nickel-based alloy wet oxidation system to safely treat caustic sulfide liquor without excessive corrosion to the materials of construction of the system. The process requires extensive analysis of the raw feed liquor and
15 gives best results with a constant composition feed.

As mentioned above, the nickel-based alloys are resistant to corrosion by caustic sulfide liquors under wet oxidation conditions, provided the pH of the liquor is maintained on the alkaline side, that is above pH 7. Likewise, the ferrous-
20 based alloys provide their best corrosion resistance at neutral or moderately alkaline pH. The wet oxidation of sulfide liquors generates acidic species which consume alkalinity. Depending on the components present, their concentration, and the pH of the caustic sulfide liquor, wet
25 oxidation may produce an oxidized liquor in which the pH is acidic, i.e. all alkalinity is consumed, and which is highly corrosive to a ferrous-based alloy wet oxidation system. To overcome this difficulty, we have devised a process which relates the species present in the caustic liquor to the

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amount of caustic required to maintain an excess of alkalinity in the liquor during wet oxidation treatment. This process allows a ferrous-based alloy wet oxidation system to safely treat caustic sulfide liquor without excessive corrosion to the materials of construction of the system provided the treatment temperature is less than about 175°C. It must be recognized that the corrosion problems need careful consideration in that the integrity of the pressurized wet oxidation system is important for both safety and economic reasons.

SUMMARY OF THE INVENTION

The invention is a process for treatment of a caustic sulfide liquor by wet oxidation in a ferrous-based alloy system which prevents excessive corrosion of the system. The process comprises the steps:

- a) analyzing the caustic sulfide liquor for initial concentrations of total alkalinity, total sulfides, mercaptans, COD, thiosulfate, total carbonate and pH to determine the amount of nonsulfidic alkalinity consumed by said liquor upon wet oxidation treatment;
- b) adding sufficient additional nonsulfidic alkalinity to said caustic sulfide liquor whereby the initial nonsulfidic alkalinity concentration plus additional nonsulfidic alkalinity concentration exceeds the nonsulfidic alkalinity consumed upon wet oxidation treatment as determined in step a; and
- c) carrying out said treatment process of wet oxidation upon said caustic sulfide liquor within said ferrous-

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based alloy system at temperatures of less than about 175°C to destroy sulfides and mercaptans and produce a treated liquor with excess nonsulfidic alkalinity concentration thereby preventing excessive corrosion to said ferrous-based alloy wet oxidation system.

BRIEF DESCRIPTION OF THE DRAWING

The Figure shows a general schematic for a wet oxidation system used to treat caustic sulfide liquors.

DESCRIPTION OF THE PREFERRED EMBODIMENTS

The Figure shows a schematic flow diagram for a wet oxidation system used for treatment of caustic sulfide scrubbing liquors. Referring to the Figure, raw caustic sulfide liquor from a storage tank 10 flows through a conduit 12 to a high pressure pump 14 which pressurizes the liquor. The raw liquor is mixed with a pressurized oxygen-containing gas, supplied by a compressor 16, within a conduit 18. The mixture flows through a heat exchanger 20 where it is heated to a temperature which initiates oxidation. The heated mixture then flows through a second heat exchanger 22 which provides auxiliary heat for startup of the system. For waste with low COD content, auxiliary heating may need to be continuously applied through the second heat exchanger 22 in order to maintain the desired operating temperature for the wet oxidation system. The heated feed mixture then enters a reactor vessel 24 which provides a residence time wherein the bulk of the oxidation reaction occurs. The oxidized liquor and oxygen depleted gas mixture then exits the reactor through

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a conduit 26 controlled by a pressure control valve 28. The hot oxidized effluent traverses the heat exchanger 20 where it is cooled against incoming raw liquor and gas mixture. The cooled effluent mixture flows through a conduit 30 to a separator vessel 32 where liquid and gases are disengaged. The liquid effluent exits the separator vessel 32 through a lower conduit 34 while the gases are vented through an upper conduit 36.

It is imperative that an excess of alkalinity be maintained throughout the ferrous-based alloy wet oxidation system when treating caustic sulfide liquor. Should the pH of the liquor drop below 7, the ferrous-based alloy of the system experiences excessive corrosion at all temperatures. Even when excess alkalinity is maintained and the liquor remains above pH 7, excessive corrosion to the ferrous-based alloy system occurs at temperatures above about 175°C. Thus the oxidation temperature of the system is maintained at less than about 175°C to maintain the integrity of the operating system. To accomplish this goal, precise analyses of certain components of the raw caustic sulfide liquor must be made. Based upon these analyses, an appropriate amount of alkalinity is added to the raw caustic sulfide liquor to maintain an excess of alkalinity throughout the wet oxidation system. It is important to avoid addition of excess alkalinity since the oxidized liquor generally must be neutralized prior to discharge to the environment. Monitoring the alkalinity or pH of the oxidized effluent with the system in operation is not effective for corrosion control since by the time the pH of the effluent is found to be acidic from insufficient

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alkalinity, the corrosion damage to the oxidation system has already occurred. The pH range for the treated liquor should be between 8 and 14.

The specific analyses required for the raw liquor include
5 pH, total alkalinity, total sulfides, mercaptide, thiosulfate, carbonate and Chemical Oxygen Demand (COD). Based on these results, the amount of nonsulfidic alkalinity consumed by the liquor upon wet oxidation treatment is calculated and compared to the nonsulfidic alkalinity available for neutralizing acid
10 formed in raw liquor. Any deficiency in available nonsulfidic alkalinity is added to the raw caustic sulfide liquor prior to wet oxidation treatment to maintain an excess of alkalinity during treatment.

The total alkalinity of a caustic solution is the sum of
15 all the titratable bases present. Total alkalinity is measured by determining the number of equivalents of strong acid required to neutralize a given volume of caustic solution to a pH of 4.5. For instance, titration of 100 milliliters of a caustic solution to a pH of 4.5 required 25 milliliters of
20 0.5 normal hydrochloric acid (0.5 equivalents/liter). The total alkalinity of the caustic solution is therefore 0.125 equivalents/liter. The total alkalinity of the liquor results not only from metal hydroxides (NaOH), carbonates (Na_2CO_3) and bicarbonates (NaHCO_3), herein termed nonsulfidic alkalinity,
25 but also from metal sulfides (NaS) and bisulfides (NaHS), as well as metal mercaptides (NaSR), herein termed sulfidic alkalinity. The total alkalinity of the caustic solution can be increased by adding basic substances, such as alkali metal hydroxides, carbonates or bicarbonates, to the solution.

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Analysis of the caustic solution for pH, total sulfides, mercaptides, thiosulfate, carbonate and COD is by well known analytical procedures. The procedures outlined by J.W. McCall in Chemical Analysis of Industrial Water, pages 125-153, are particularly useful. The measurement of pH is by a glass electrode against a standard calomel electrode. Total sulfides and mercaptide concentrations are determined by potentiometric titration with standardized silver nitrate solution. Thiosulfate is analyzed by an iodine titration method. Carbonate may be measured either by barium carbonate precipitation followed by acidification and back-titration with standardized base or by a dissolved inorganic carbon procedure. COD is measured by the silver catalyzed dichromate oxidation method, Standard Methods No. 5220.

Considering the alkalinity requirement for each of the components, we can write the following oxidation equations:

Sulfidic Components



Organic Components



These equations summarize the alkalinity requirements for the various species present in caustic sulfide liquors.

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Sulfide, when present as sodium sulfide, does not consume any alkalinity when it is oxidized. Sulfide, when present as bisulfide, NaHS, consumes one equivalent of alkalinity for each mole of NaHS oxidized. Each mole of thiosulfate oxidized
5 consumes two equivalents of alkalinity. Sulfur, when present as sodium mercaptide, NaSR, consumes two equivalents of alkalinity for each mole of NaSR-sulfur oxidized.

The organic portion of the Chemical Oxygen Demand (COD) consumes one equivalent of alkalinity for each mole of oxygen
10 used in the oxidation of organic species. The organic portion of the COD is calculated by measuring the total COD of the caustic waste and subtracting the sulfidic COD which is calculated from the analytical concentrations determined for each sulfidic component. That is:

$$15 \quad \text{COD (total)} - \text{COD (sulfidic)} = \text{COD (organic)}$$

The amount of COD (organic) which is removed will depend upon oxidation conditions and will vary from about 5 percent at 110°C to about 30 percent at 175°C. Thus the alkalinity required will be some fraction of the total COD (organic),
20 depending on operating temperature.

$$\text{COD (organic)} \times (5\% \text{ to } 30\%) = \text{COD (organic removed)}$$

In calculating the amount of alkalinity consumed by the liquor upon wet oxidation treatment, determining the concentrations of sulfide, $\{S^{2-}\}$, and bisulfide, $\{HS^{-}\}$, is
25 important since sulfide consumes no alkalinity upon oxidation

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while bisulfide consumes one equivalent of alkalinity per mole oxidized. Likewise, an accurate determination of solution pH is important since pH affects the dissociation of hydrogen sulfide, sodium bisulfide and sodium sulfide.

5 As mentioned above, the pH of the caustic liquor is determined with a pH meter fitted with a special glass electrode for measurement of pH in the range 11 to 14 against a saturated calomel electrode. The observed pH must be corrected for the sodium ion concentration in the liquor by
10 using a nomograph chart supplied with the glass electrode.

The total sulfides concentration C_s of the liquor is determined by potentiometric titration with standard silver nitrate solution. This total sulfides concentration plus the accurately determined pH value allows calculation of the
15 sulfide concentration, $\{S^{2-}\}$, and bisulfide concentration, $\{HS^{-}\}$, as follows:

The dissociation constants for hydrogen sulfide (H_2S) are:
 $K_1 = 5.7 \times 10^{-8}$ and $K_2 = 1.2 \times 10^{-15}$
and $\{H^{+}\} = 10^{-pH}$

20 According to acid-base equilibrium principles, it follows that:

$$\{HS^{-}\} = C_s [K_1 \{H^{+}\} / \{H^{+}\}^2 + K_1 \{H^{+}\} + K_1 K_2]$$

$$\{S^{2-}\} = C_s [K_1 K_2 / \{H^{+}\}^2 + K_1 \{H^{+}\} + K_1 K_2]$$

Using these equations the ratio of concentrations of bisulfide
25 to sulfide can be calculated over the pH range 9 to 14. This

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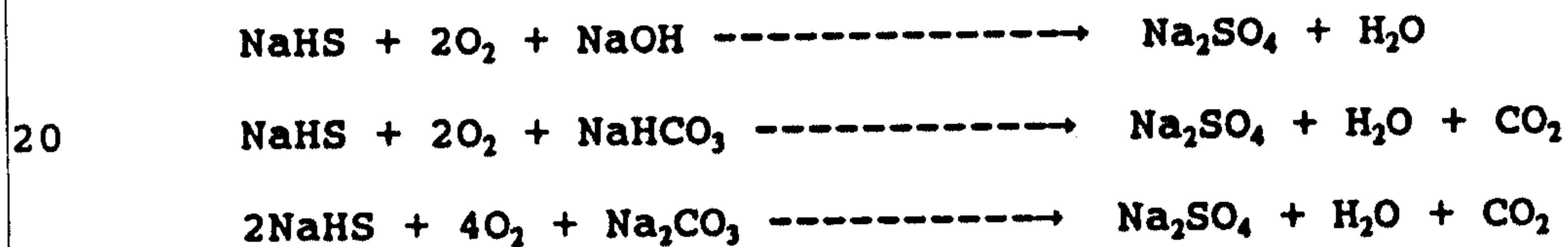
ratio, from the above two equations, simplifies to $\{HS^-\}/\{S^{2-}\}$
 $= K_1\{H^+\}/K_1K_2$ which gives the following:

	<u>pH</u>	<u>$\{H^+\}$</u>	<u>$\{HS^-\}/\{S^{2-}\}$</u>
	9	1×10^{-9}	833,333.
5	10	1×10^{-10}	83,333.
	11	1×10^{-11}	8,333.
	12	1×10^{-12}	833.
	13	1×10^{-13}	83.3
	13.5	3.16×10^{-14}	26.4
10	14	1×10^{-14}	8.33

With the total sulfides concentration C, determined and the bisulfide to sulfide ratio determined from the pH of the liquor, the individual sulfide species concentrations are easily determined algebraically.

15 From these calculations it can be seen that even at pH 14 the vast majority of sulfide is present as bisulfide, HS^- , which consumes one equivalent of alkalinity upon oxidation.

Nonsulfidic Alkalinity Components



The amount of nonsulfidic alkalinity present in the liquor and available for consumption by the sulfidic species also requires determination. The reactions shown above for

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nonsulfidic alkalinity components show that each mole of sodium hydroxide and sodium bicarbonate provides one equivalent of alkalinity while each mole of sodium carbonate provides two equivalents of alkalinity. The pH of the liquor
5 also affects the proportion of carbonate to bicarbonate present in the liquor.

As mentioned above, the total carbonate concentration of the liquor may be measured by the barium carbonate precipitation procedure or by a dissolved inorganic carbon
10 procedure. The chosen analysis provides a total carbonate concentration, $C_{\text{carbonate}}$. This value plus the accurately determined pH value, determined as described earlier, allows calculation of the carbonate concentration, $\{\text{CO}_3^{2-}\}$, and bicarbonate concentration, $\{\text{HCO}_3^-\}$, as follows:

15 The dissociation constants for carbonic acid (H_2CO_3) are:

$$K_1 = 4.3 \times 10^{-7} \quad \text{and} \quad K_2 = 5.6 \times 10^{-11}$$

$$\text{and} \quad \{\text{H}^+\} = 10^{-\text{pH}}$$

According to acid-base equilibrium principles, it follows that:

$$20 \quad \{\text{HCO}_3^-\} = C_{\text{carbonate}} \left[\frac{K_1 \{\text{H}^+\}}{\{\text{H}^+\}^2 + K_1 \{\text{H}^+\} + K_1 K_2} \right]$$

$$\{\text{CO}_3^{2-}\} = C_{\text{carbonate}} \left[\frac{K_1 K_2}{\{\text{H}^+\}^2 + K_1 \{\text{H}^+\} + K_1 K_2} \right]$$

Using these equations the ratio of concentrations of bicarbonate to carbonate can be calculated over the pH range 8 to 14. This ratio, from the above two equations, simplifies
25 to $\{\text{HCO}_3^-\}/\{\text{CO}_3^{2-}\} = K_1 \{\text{H}^+\}/K_1 K_2$ which gives the following:

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	pH	$\{H^+\}$	$\{HCO_3^-\}/\{CO_3^{2-}\}$
	8	1×10^{-8}	205.7
	9	1×10^{-9}	20.6
	10	1×10^{-10}	2.06
5	11	1×10^{-11}	0.206
	12	1×10^{-12}	0.0206
	13	1×10^{-13}	0.00206
	14	1×10^{-14}	0.000206

With the total carbonate concentration $C_{\text{carbonate}}$ determined and
 10 the bicarbonate to carbonate ratio determined from the pH of
 the liquor, the individual carbonate species concentrations
 are easily determined algebraically.

This leaves the concentration of sodium hydroxide to be
 determined. An extensive examination of the interaction of
 15 pH, hydroxide, sulfide, mercaptan and carbonate is set forth
 in the J.W. McCoy reference at pages 138-153. The hydroxide
 concentration for various conditions is obtained from the
 accurate pH measurement and previously determined species.
 These conditions are:

20 Condition I. pH > 13.0

$$\{NaOH\} = T. \text{ Alkalinity} - [(\{C_s\} + \{RSNa\} + \{CO_3^{2-}\}) - \{HS^-\}]$$

Condition II. pH 12.0-13.0

$$\{NaOH\} = \{OH^-\} = 10^{-(14-pH)}$$

Condition III. pH < 12.0

25 $\{NaOH\} = 0$

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Finally, totalling up the nonsulfidic alkalinity of the liquor from carbonate, bicarbonate and sodium hydroxide and comparing this total with the alkalinity requirements for sulfide, bisulfide, thiosulfate, mercaptan sulfur and organic
5 COD determines if sufficient alkalinity is present to maintain the liquor in the proper operating pH range to prevent excessive corrosion to the wet oxidation system. If there is insufficient nonsulfidic alkalinity to satisfy the requirements of the alkalinity consuming species, then
10 additional alkalinity must be added to the liquor to prevent corrosion of the wet oxidation system. In the ferrous-based alloy wet oxidation systems, the oxidation temperature is maintained at less than about 175°C to prevent corrosion even when excess alkalinity is maintained in the treated caustic
15 liquor.

EXAMPLE

A caustic sulfide scrubbing liquor with a COD of 13.5 g/l and pH 12.78 was treated in a laboratory autoclave with excess oxygen by wet oxidation at 200°C for 60 minutes. The oxidized
20 liquor had a COD of 0.9 g/l and pH 1.5. This pH would cause excessive corrosion of a ferrous-based alloy wet oxidation system at this temperature, or even at temperatures as low as 100°C.

Further analyses of the raw liquor for components of
25 interest gave the following results:

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	pH	12.78
	COD	13.5 g/l
	T. Alkalinity	0.46 eq./l
	T. Sulfides	6.8 g/l
5	Thiosulfate	N.D.
	Mercaptan Sulfur	N.D.
	T. Carbonate	2.3 g/l

The oxygen demand for 6.8 g/l total sulfides is calculated to be 13.6 g/l, thus essentially all the 13.5 g/l of measured COD is from sulfides. No mercaptan sulfur or thiosulfate was detected. At pH 12.78 the ratio $\{HS^-\}/\{S^{2-}\} = 138.3$. Thus, essentially all sulfur is in the bisulfide form which on oxidation consumes one equivalent alkalinity per mole. The 6.8 g/l sulfide calculates to 7.013 g/l bisulfide which is 0.2125 eq./l.

The total carbonate concentration was 2.3 g/l as sodium carbonate. At pH 12.78, the ratio $\{HCO_3^-\}/\{CO_3^{2-}\} = 0.0034$. Thus, essentially all inorganic carbon is present in the carbonate form. The 2.3 g/l sodium carbonate provides 0.0434 eq./l alkalinity.

Since the liquor has pH 12.78, Condition II occurs and $\{NaOH\} = \{OH^-\}$ which is $10^{-(14-12.78)} = 10^{-1.22}$ or 0.0603 moles/l. Thus $\{NaOH\} = 0.0603$ eq./l. The available nonsulfidic alkalinity is the sum of carbonate and hydroxide which totals 0.1037 eq./l, about half of the 0.2125 eq./l alkalinity consuming bisulfide. At least 0.1125 eq./l alkalinity must be added to this liquor to prevent excessive corrosion and allow safe operation of a ferrous-based alloy wet oxidation system at temperatures less than about 175°C.

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Another caustic sulfide scrubbing liquor was analyzed for components of interest with the following results:

	pH	13.79
	COD	73.6 g/l
5	T. Alkalinity	2.28 eq./l
	T. Sulfides	47.7 g/l
	Thiosulfate	N.D.
	Mercaptan Sulfur	N.D.
	T. Carbonate	6.63 g/l

10 The oxygen demand for 47.7 g/l total sulfides is calculated to be about 95 g/l, thus all the measured COD is from sulfides. No mercaptan sulfur or thiosulfate was detected. At pH 13.79 the ratio $\{HS\}/\{S^{2-}\} = 13.52$ which results in a distribution of 93% bisulfide and 7% sulfide.

15 The 47.7 g/l total sulfides is divided into 44.36 g/l bisulfide and 3.34 g/l sulfide. The 44.36 g/l bisulfide consumes 1.386 eq./l alkalinity upon oxidation.

At pH 13.79 the ratio $\{HCO_3\}/\{CO_3^{2-}\} = 0.0003$ so essentially all inorganic carbon is in the carbonate form

20 which is 0.125 eq./l. At this pH the sodium hydroxide concentration is calculated as for Condition I:

$$\begin{aligned} \{NaOH\} &= T. Alkalinity - [\{C_t\} + \{RSNa\} + \{CO_3^{2-}\}] - \{HS\} \\ \{NaOH\} &= 2.28 \text{ eq./l} - [(1.49 + 0 + 0.125) - 1.386] \text{ eq./l} \\ \{NaOH\} &= 2.28 - 0.229 = 2.051 \text{ eq./l} \end{aligned}$$

25 Thus the available nonsulfidic alkalinity is 2.175 eq./l while the bisulfide consumes only 1.386 eq./l alkalinity on

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oxidation. To verify these results wet oxidation of this raw liquor at 200°C for 60 minutes in a laboratory autoclave resulted in an oxidized liquor with a COD of 1.7 g/l and pH 9.89. Thus, the liquor requires no added alkalinity for safe
5 treatment in a ferrous-based alloy wet oxidation system at temperatures less than about 175°C.

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THE EMBODIMENTS OF THE INVENTION IN WHICH AN EXCLUSIVE PROPERTY OR PRIVILEGE IS CLAIMED ARE DEFINED AS FOLLOWS:

1. A process for treatment of caustic sulfide liquor by wet oxidation in a ferrous-based alloy system comprising the steps:

a) analyzing the caustic sulfide liquor to determine the initial concentrations therein of total alkalinity, total sulfides, mercaptans, COD, thiosulfate, total carbonate and pH, and determining from the results of this analysis the amount of nonsulfidic alkalinity consumed by said liquor upon wet oxidation treatment and the amount of nonsulfidic alkalinity initially available in said liquor;

b) adding a sufficient amount of additional nonsulfidic alkalinity to said caustic sulfide liquor such that the sum of the initially available nonsulfidic alkalinity and said additional nonsulfidic alkalinity exceeds the amount of nonsulfidic alkalinity consumed upon wet oxidation treatment as determined in step a; and

c) carrying out said treatment process of wet oxidation upon said caustic sulfide liquor within said ferrous-based alloy wet oxidation system at temperatures less than about 175°C to destroy sulfides and mercaptans and produce a treated liquor with excess nonsulfidic alkalinity.

2. A process according to claim 1 wherein said caustic sulfide liquor is produced from a scrubbing process.

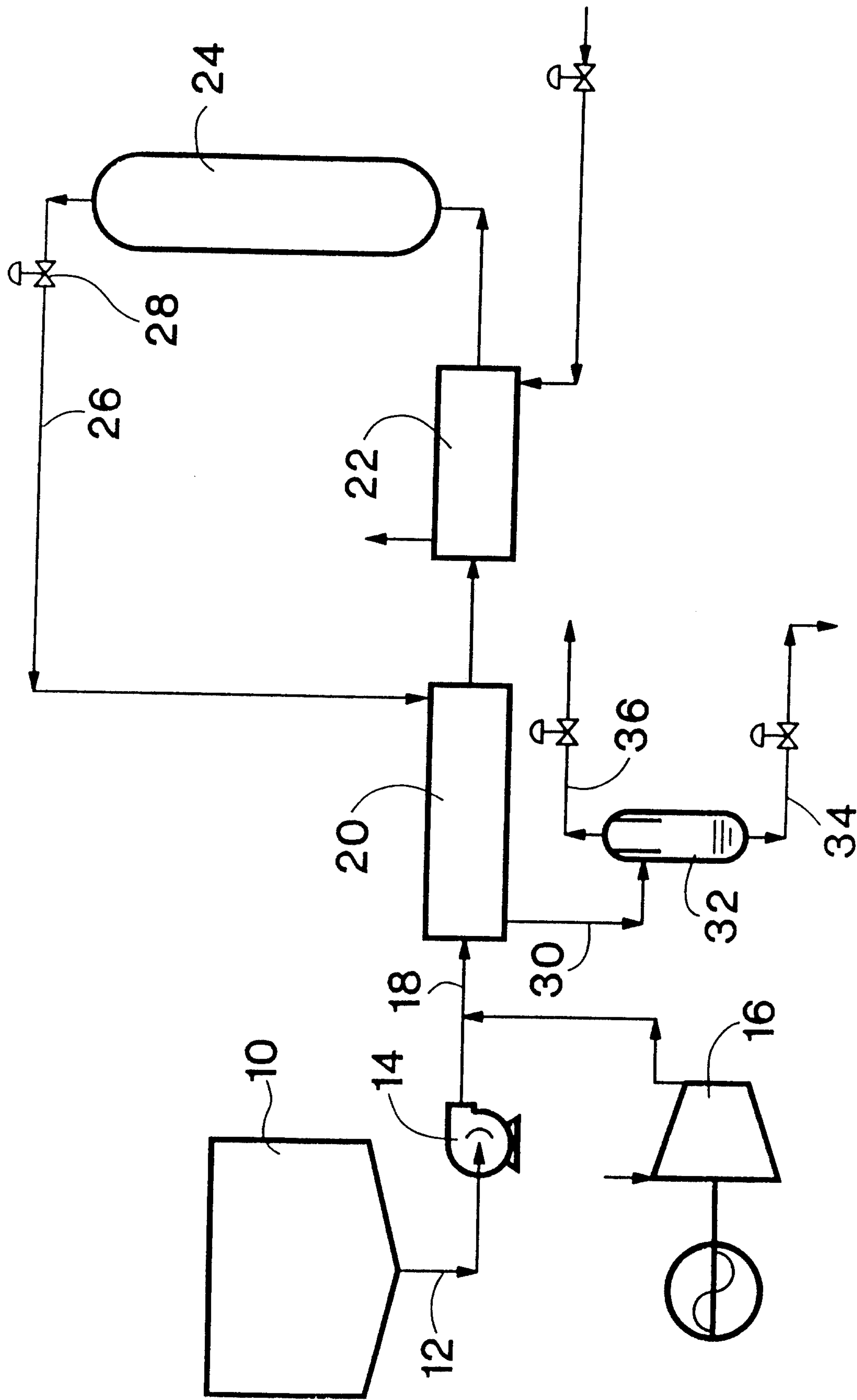
3. A process according to claim 1 wherein said caustic sulfide liquor is produced from an extraction process.

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4. A process according to claim 1 wherein said additional nonsulfidic alkalinity is added to said liquor in the form of alkali metal hydroxide.

5. A process according to claim 1 wherein said additional nonsulfidic alkalinity is added to said liquor in the form of alkali metal carbonate or bicarbonate.

6. A process according to claim 1 wherein the pH range for said treated liquor is between 8 and 14.



FIGURE



