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(54) Title: TRIANIONIC PINCER LIGANDS, A CR(III)/CR(V) CATALYTIC SYSTEM AND ITS USE FOR CATALYTIC AEROBIC OXIDATION OF ORGANIC SUBSTRATES

(57) Abstract: An embodiment of the invention is a novel Cr(V) OCO³⁻ trianionic pincer ligand complex. Another embodiment of the invention is a catalytic method for oxidation of a substrate aerobically in the presence of a source of oxygen, where the novel Cr(V) OCO³⁻ trianionic pincer ligand complex acts as the catalyst. The substrate can be a phosphine, amine, sulfide, alkene, alkane or a second metal complex. Another embodiment of the invention is directed to NCN pincer ligands that can form trianionic pincer ligand complexes.



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DESCRIPTION

TRIANIONIC PINCER LIGANDS, A CR(III)/CR(V) CATALYTIC SYSTEM AND ITS
USE FOR CATALYTIC AEROBIC OXIDATION OF ORGANIC SUBSTRATES

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The subject invention was made with government support under a research project supported by National Science Foundation, Contract No. CHE-0748408. The government has certain rights to this invention.

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CROSS-REFERENCE TO A RELATED APPLICATION

This application claims the benefit of U.S. Provisional Application Serial No. 61/156,946, filed March 3, 2009, the disclosure of which is hereby incorporated by reference in its entirety, including all figures, tables and drawings.

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BACKGROUND OF THE INVENTION

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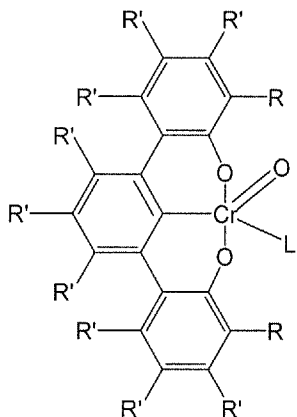
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Transient, high oxidation state, metal-oxo species are key intermediates in life sustaining biological energy conversion transformations, such as water oxidation at the oxygen-evolving complex (OEC) in photosynthesis II, and cellular substrate oxidation at heme iron in cytochrome P450s, peroxidases, and catalases. Despite the opposite function of the two systems (O_2 cleavage; P450, O_2 formation; PS II), the principle design feature involves managing oxygen atom transfer (OAT) from highly reactive metal-oxo fragments. In each case, a unique ligand supports an otherwise unstable high oxidation state metal-oxo species. In the oxygen evolving complex of PS II, clustered multiple Mn centers of the S4 Kok state ($Mn(IV)_3Mn(V)$) act as elaborate ligands for a manganyl-oxo intermediate. For P450, heme radical cation formation enables ferryl $Fe(IV)=O$ to exist. Not all high oxidation state metal-oxo species are productive for living systems. For example, Cr(VI) is a carcinogen and causes cellular oxidative DNA damage. However, Cr(VI) is not an active oxidant, instead, ascorbic or glutathione reduction to Cr(III) leads to formation of Cr(IV) and Cr(V) intermediates.

Generating reactive intermediates incurs negative outcomes even nature cannot avoid. Oxidative enzyme degradation pathways are inevitable but protein repair processes are built-in. In artificial systems, the catalytic cycle ends once the ligand degrades, thus limiting turnover. Much of the challenge in developing OAT catalytic systems is to obtain robust, degradation impervious catalysts. To control the geometry and electronic properties of metal ions, chemists create customized ligands to mimic nature. Porphyrin, corrole, catechol, 2-hydroxyacid, and salen ligands can stabilize Cr(V), including a few structurally characterized examples. Some Cr(V)=O oxo complexes are known that participate as catalysts in OAT to sulfides, phosphines, olefins, alkynes, and enantioselective adaptations but are relatively unstable. A stable Cr complex would permit OAT in a commercially viable manner. Hence, there is a need for custom ligands, catalytic systems and their use in catalytic oxidation of organic compounds.

BRIEF SUMMARY OF THE INVENTION

Embodiments of the invention are directed to a Cr(V) OCO^{3-} trianionic pincer ligand complex of the structure:



where: R is independently a C3 to C12 alkyl or aryl group; R' is independently a H, C₁ to C₁₂ alkyl or aryl group; and L is an oxygen comprising ligand. For example, the pincer ligand with R = *t*-butyl, all R' = H and L = tetrahydrofuran (THF) is a useful for the Cr(V) OCO³⁻ trianionic pincer ligand complex.

Embodiments of the invention are directed to NCN^{3-} ligands that can be used to form trianionic pincer ligand complexes with Cr and other metals. These NCN ligands permit complexes such as that with the above OCO ligands, where a pair of six-member rings or five-member rings includes the NCN anionic sites and metal ion.

Another embodiment of the invention is a catalytic method for oxidation of a substrate that employs the Cr(V) OCO^{3-} trianionic pincer ligand complex as catalyst. A substrate is oxidized by the Cr(V) OCO^{3-} trianionic pincer ligand complex with reduction of the Cr(V) to Cr(III), which is oxidized back to the Cr(V) OCO^{3-} trianionic pincer ligand complex by molecular oxygen. The oxidation can be carried out homogeneously in a solvent such as an aliphatic hydrocarbon, aromatic hydrocarbon, or ether. The oxygen source can be pure oxygen or oxygen diluted in an inert gas. The oxygen source can be air. Various substrates can be oxidized including phosphines, amines, sulfides, alkenes, alkanes or other metal complexes.

Another embodiment of the invention is directed to the preparation of Cr(V) OCO^{3-} trianionic pincer ligand complexes. In this method, an OCOH_3 pincer ligand is converted to an OCOHM_2 pincer complex, the metal ion of the OCOHM_2 pincer complex is exchanged with a Cr(III) salt, and the Cr(III) is oxidized to Cr(V). The OCOH_3 pincer ligand can be $(^t\text{BuOCO})\text{H}_3$, which can be converted to an OCOHK_2 pincer complex by reaction with KH. The OCOHK_2 pincer complex can undergo metal exchange with the salt $\text{CrCl}_2\text{Me}(\text{THF})_3$ to yield a Cr(III) OCO^{3-} trianionic pincer ligand complex. The Cr(III) OCO^{3-} trianionic pincer ligand complex can be oxidized by oxygen to the desired Cr(V) OCO^{3-} trianionic pincer ligand complex.

BRIEF DESCRIPTION OF THE DRAWINGS

Figure 1 is a reaction scheme of a catalytic cycle for the oxidation of a substrate using a Cr(V) OCO^{3-} trianionic pincer ligand complex **7** according to an embodiment of the invention.

Figure 2 is a reaction scheme for the transformation of a pincer ligand $(^t\text{BuOCO})\text{H}_3$ **1** to dipotassium-3,3'-di-*tert*-butyl-2,2'-dimethoxy-1,1':3,1'-terphenoxide, $(^t\text{BuOCO})\text{HK}_2(\text{THF})$ **9**.

Figure 3 is a reaction scheme for the transformation of complex $(^t\text{BuOCO})\text{HK}_2(\text{THF})$ **9** to a Cr(III) OCO^{3-} trianionic pincer ligand complex **8**.

Figure 4 is a reaction scheme for the transformation of a Cr(III) OCO^{3-} trianionic pincer ligand complex **9** to a Cr(V) OCO^{3-} trianionic pincer ligand complex **8** according to an embodiment of the invention.

Figure 5 is: **A** an ORTEP representation of $(t\text{BuOCO})\text{HK}_2(\text{THF})$ (**10**) with ellipsoids drawn at the 50% probability level (H-atoms and a THF solvent molecule bound to K1 are omitted for clarity; and **B** Molecular structure of **10** depicting K- π -interactions ($t\text{Bu}$ groups, H-atoms and THF omitted for clarity).

Figure 6 is a ^1H NMR spectrum of $[t\text{BuOCO}]\text{Cr}^{\text{III}}(\text{THF})_3$ (**9**) displaying broad paramagnetically shifted resonances.

Figure 7 is an ORTEP representation of $[t\text{BuOCO}]\text{Cr}^{\text{III}}(\text{THF})(\text{DME})$ (**9-DME**) with ellipsoids drawn at the 50% probability level and hydrogen atoms removed for clarity.

Figure 8 is an infrared spectrum of Cr(V) $[t\text{BuOCO}]\text{Cr}^{\text{V}}=\text{O}(\text{THF})$ (**8**) according to an embodiment of the invention.

Figure 9 is a reaction scheme for the preparation of NCN ligand **2** - R = mesityl according to an embodiment of the invention.

Figure 10 is a ^1H NMR spectrum of (**2** - R = mesityl) in C_6D_6 .

Figure 11 is a $^{13}\text{C}\{^1\text{H}\}$ NMR spectrum of (**2** R = mesityl) in C_6D_6 .

Figure 12 is a reaction scheme for the preparation of NCN ligand **3** - R = 3,5-bis-(trifluoromethyl)phenyl according to an embodiment of the invention.

Figure 13 is a ^1H NMR spectrum of (**3** - R = 3,5-bis-(trifluoromethyl)phenyl) in C_6D_6 .

Figure 14 is a $^{13}\text{C}\{^1\text{H}\}$ NMR spectrum of (**3** - R = 3,5-bis-(trifluoromethyl)phenyl) in C_6D_6 .

Figure 15 is an Ortep drawing of the molecular structure of (**3** - R = 3,5-bis-(trifluoromethyl)phenyl) with ellipsoids presented at the 50% probability level.

Figure 16 is a reaction scheme for the preparation of NCN ligand **4** - R = 3,5-bis-(trifluoromethyl)phenyl according to an embodiment of the invention.

Figure 17 is a ^1H NMR spectrum of (**4** - R = 3,5-bis-(trifluoromethyl)phenyl) in C_6D_6 .

Figure 18 is a $^{13}\text{C}\{^1\text{H}\}$ NMR spectrum of (**4** - R = 3,5-bis-(trifluoromethyl)phenyl) in C_6D_6 .

Figure 19 is a reaction scheme for the preparation of NCN ligand **5** – $R = Si^iPr_3$ according to an embodiment of the invention.

Figure 20 is a ^1H NMR spectrum of (**5** – $\text{R} = \text{Si}^i\text{Pr}_3$) in C_6D_6 .

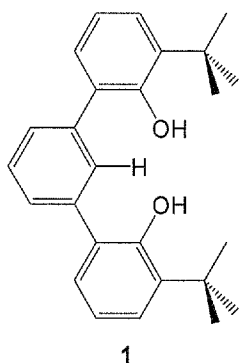
Figure 21 is a $^{13}\text{C}\{^1\text{H}\}$ NMR spectrum of (**5** – **R** = Si^iPr_3) in C_6D_6 .

Figure 22 is a reaction scheme for the preparation of the Cr NCN trianionic pincer ligand $\text{NCNCr}^{\text{III}}(\text{THF})_3$ **11** according to an embodiment of the invention.

Figure 23 is a ^1H NMR spectrum of $\text{NCNCr}^{\text{III}}(\text{THF})_3$ **11** in C_6D_6 .

DETAILED DESCRIPTION OF THE INVENTION

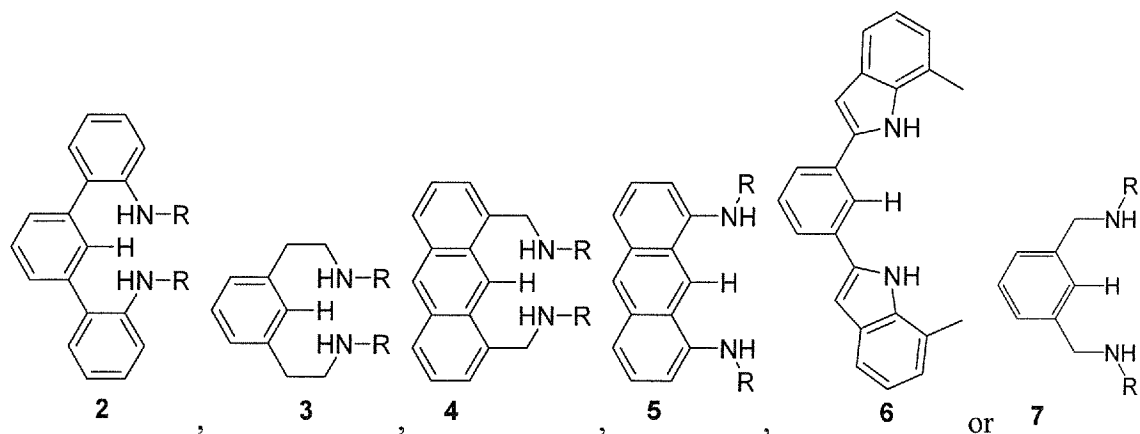
Embodiments of the invention are directed to Cr(V) complexes stabilized by an
trianionic pincer ligand. A novel Cr(V) OCO³⁻ trianionic pincer ligand complex
comprises a trianionic pincer ligand of the structure (^tBuOCO)H₃ (**1**):



Other OCO³⁻ trianionic pincer ligand can be employed in embodiments of the invention. All carbons in **1** at positions 3,4,5,4',5',6',3'',4'' and 5'' can be independently substituted, for example, with alkyl groups such as methyl, ethyl, *n*-propyl, *i*-propyl, *n*-butyl, *i*-butyl *t*-butyl, or larger alkyl groups or any other substituent in a manner that does not inhibit formation of the metal complex, as is obvious to those skilled in the art. For example, large substitution at the two ortho positions to the bonds linking adjacent aryl rings can inhibit co-planarity of the adjacent aromatic rings and inhibit chelation of a metal ion. Other sterically hindering alkyl groups can be substituted for the *t*-butyl groups of **1**.

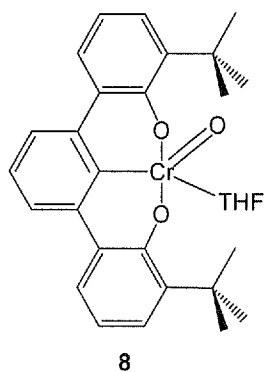
Other embodiments of the invention are directed to NCN pincer ligands that can form NCN^{3-} complexes with Cr(V), Cr(III) or other metal ions. These NCN pincer

ligands form complexes analogous to those formed by **1** where the NCN and metal ions form a pair of six-member rings or a pair of five-member rings in the complex. The NCN pincer ligands are:



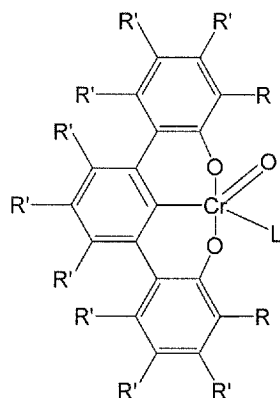
where R is 2,6-bis-(*i*-propyl)phenyl, 3,5-bis-(methyl)phenyl, 3,5-bis-(trifluoromethyl)phenyl, 3,5-bis-(*i*-propyl)phenyl, mesityl, or tri-*i*-propylsilyl. Again, as with ligand **1** all carbons positions not shown with an H substituent or an N substituent can be independently substituted, for example, with alkyl groups such as methyl, ethyl, *n*-propyl, *i*-propyl, *n*-butyl, *i*-butyl *t*-butyl, or larger alkyl groups or any other substituent in a manner that does not inhibit formation of the metal complex, as is obvious to those skilled in the art.

Compound **1** has the ability to replace the three hydrogens shown in the illustrating structure with bonds to a Cr ion for various oxidation states of chromium. The conjugated system remains rigid, forming a stable complex. A Cr(V) OCO³⁻ trianionic pincer ligand complex according to an embodiment of the invention has the structure:



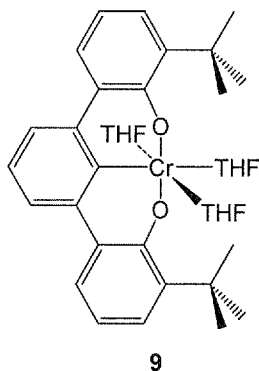
Other Cr(V) OCO³⁻ trianionic pincer ligand complexes according to embodiments of the invention can have OCO³⁻ trianionic pincer ligands where any carbon in the ligand of **8** can be independently substituted at positions 3,4,5,4',5',6',3'',4'' and 5'', for example,

with alkyl groups such as methyl, ethyl, *n*-propyl, *i*-propyl, *n*-butyl, *i*-butyl *t*-butyl, or larger alkyl groups or any other substituent that does not inhibit formation of the metal complex, as is obvious to those skilled in the art. Other sterically hindering alkyl groups can be substituted for the *t*-butyl groups of **8**. Other complexing solvent molecules can be substituted for tetrahydrofuran (THF), for example other cyclic ethers. The complex can have the structure:



where: R is independently a C3 to C12 alkyl or aryl group; R' is independently a H, C1 to C12 alkyl or aryl group; and L is any neutral coordinating ligand.

The Cr(V) OCO^{3-} trianionic pincer ligand complexes according to embodiments of the invention can be used in a method of aerobic oxidation of a substrate. The method is illustrated in Figure 1, where OAT occurs between complex **8** and a Cr(III) OCO^{3-} trianionic pincer ligand complex **9** of the structure:



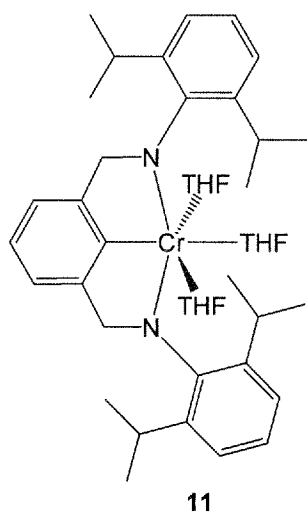
Other Cr(V) OCO^{3-} trianionic pincer ligand complexes according to embodiments of the invention can have OCO^{3-} trianionic pincer ligands where any carbon in the ligand of **9** can be independently substituted at positions 3,4,5,4',5',6',3'',4'' and 5'', for example, with alkyl groups such as methyl, ethyl, *n*-propyl, *i*-propyl, *n*-butyl, *i*-butyl *t*-butyl, or larger alkyl groups or any other substituent that does not inhibit formation of the metal complex, as is obvious to those skilled in the art. Other sterically hindering alkyl groups

can be substituted for the *t*-butyl groups of **9**. Other complexing solvent molecules can be substituted for tetrahydrofuran (THF), for example other cyclic ethers.

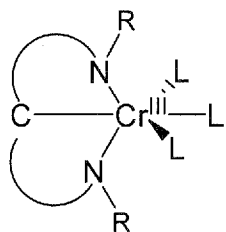
The aerobic oxidation can use oxygen from air or any stream comprising oxygen and any gas or combination of gases that are inert towards **8** and **9**. As can be seen in Figure 1, the Cr(V) OCO³⁻ trianionic pincer ligand complex **8** reacts with a substrate (S) that oxidizes to an oxidized substrate (SO) with the formation of complex **9**. The substrate can be a phosphine, amine, sulfide, alkene, alkane or another metal complex. For example: a phosphine, such as triphenyl phosphine, can be oxidized to a phosphine oxide, such as triphenyl phosphine oxide; an amine can be oxidized to an amine oxide; a sulfide can be oxidized to a sulfoxide; an alkene can be oxidized to an epoxy compound; or an active C-H bond of an alkane functionality can be oxidized to an alcohol under homogeneous conditions. The Cr(III) OCO³⁻ trianionic pincer ligand complex **9** is transformed by oxygen back to the Cr(V) OCO³⁻ trianionic pincer ligand complex **8** in the catalytic cycle.

An embodiment of the invention is a method for preparing the Cr(V) OCO³⁻ trianionic pincer ligand complex **8**. The pincer complex **8** is stable under normal ambient conditions. The catalyst preparation proceeds from (*t*BuOCO)H₃ (**1**) or its substituted equivalent by reaction with a metal hydride, for example an alkali metal hydride, for example potassium hydride in THF or other weakly complexing polar aprotic solvents. The transformation from **1** with KH is illustrated in Figure 2 to yield a dipotassium-3,3"-di-*tert*-butyl-2,2"-dimethoxy-1,1':3,1"-terphenoxide, (*t*BuOCO)HK₂(THF) **10**. A metal exchange of the alkali metal and chromium can be carried out by double salt metathesis/C-H bond activation between **10** and as shown in Figure 3 to yield **9**. The oxidation of **9** to **8** is shown in Figure 4 where **9** is oxidized by oxygen in a toluene solution. The catalytic cycle for oxidation of a substrate (S) in Figure 1, includes an equivalent oxidation of **9** to **8**.

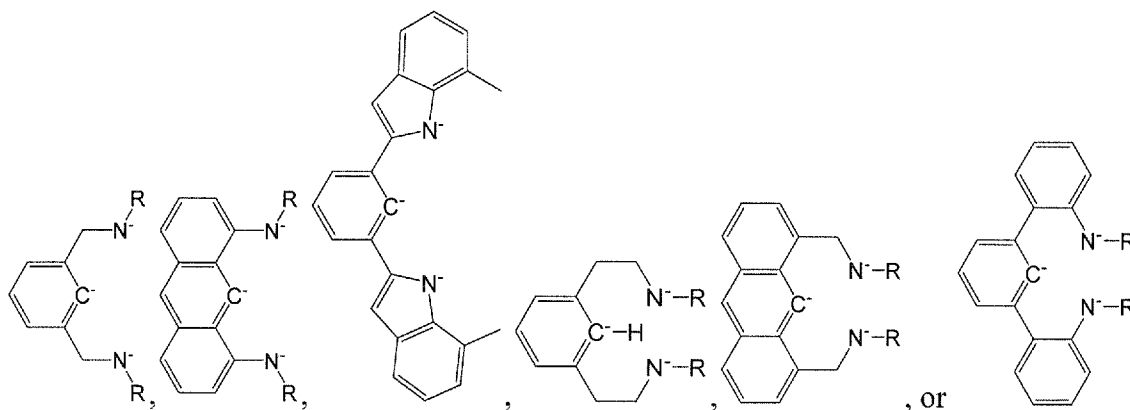
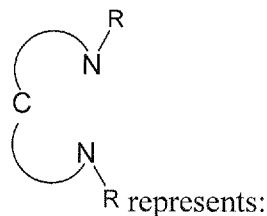
Using an NCN pincer ligand, for example with compound **7** above, a Cr(III) trianionic pincer ligand complex, for example the complex:



that has the NCN trianionic pincer ligand from **7**. In like manner to the Cr complex having two five-member rings from **7**, complexes from **5**, and **6** can be formed. Furthermore, in like manner, Cr complexes having six-membered rings can be formed starting with **2**, **3**, and **4**. Such Cr(III) NCN trianionic pincer complex has the general structure:



where the portion:



where R is 2,6-bis-(*i*-propyl)phenyl, 3,5-bis-(methyl)phenyl, 3,5-bis-(trifluoromethyl)phenyl, 3,5-bis-(*i*-propyl)phenyl, mesityl, or tri-*i*-propylsilyl and L is any neutral coordinating ligand. Neutral coordinating ligands can be ethers, for example tetrahydrofuran. The neutral ligands can have an oxygen group.

MATERIALS AND METHODS

Unless specified otherwise, all manipulations were performed under an inert atmosphere using standard Schlenk or glovebox techniques. Pentane, hexanes, toluene, diethyl ether (Et₂O), tetrahydrofuran (THF), and 1,2-dimethoxyethane (DME) were dried using a GlassContour drying column. Benzene-*d*₆ (Cambridge Isotopes) was dried over sodium-benzophenone ketyl, distilled or vacuum transferred, and stored over 4 Å molecular sieves. CrCl₂Me(THF)₃ was prepared according to published procedures. All other reagents were purchased from commercial vendors and used without further purification. NMR spectra were obtained on Varian Gemini 300 MHz, Varian Mercury Broad Band 300 MHz, or Varian Mercury 300 MHz spectrometers. Chemical shifts are reported in δ (ppm). For ¹H and ¹³C NMR spectra, the solvent peak was referenced as an internal reference. Infrared spectra were obtained on a Thermo scientific Nicolet 6700 FT-IR. Combustion analyses were performed at Complete Analysis Laboratory Inc., Parsippany, New Jersey.

Synthesis of dipotassium-3,3''-di-*tert*-butyl-2,2''-dimethoxy-1,1':3,1''-terphenoxide (10). In a nitrogen filled glovebox, 597 mg (1.59 mmol) of **1** was dissolved in 5 mL of THF. In a separate vial 128 mg (2.01 eq, 3.20 mmol) of potassium hydride was suspended in 2 mL of THF. The solution containing **1** was added to the potassium hydride suspension and stirred vigorously at room temperature for 4 h. The solution was then filtered and all volatiles removed in vacuo to provide a colorless oil. The oil was triturated with pentane (3 x 1 mL) to yield **10** as a white powder (705 mg, 75%). ¹H NMR (300 MHz, THF-*d*₈, δ): 7.98 ppm (s, 1H, H¹), 7.28 ppm (t, ³*J* = 7.79 Hz, 1H, H⁴), 6.96 ppm (dd, ³*J* = 7.33 Hz, ⁴*J* = 1.83 Hz 2H, H^{3,3'}), 6.89 ppm (dd, ³*J* = 7.79 Hz, ⁴*J* = 2.29 Hz, 2H, H^{8,8'}), 6.86 ppm (dd, ³*J* = 7.33 Hz, ⁴*J* = 2.29 Hz, 2H, h^{6,6'}), 5.99 ppm (dd, *J* = ³7.33, ³*J* = 7.33 Hz, 2H, H^{7,7'}), 1.45 ppm (s, 18H, H^{12,12'}). ¹³C NMR (75 MHz, THF-*d*₈, δ): 169.6 ppm (C^{10,10'}), 144.7 ppm (C^{2,2'}), 137.2 ppm (C^{9,9'}), 134.5 ppm (C¹), 132.19 ppm (C^{5,5'}), 128.55 ppm (C⁴), 128.07 ppm (C^{6,6'}), 126.43 ppm (C^{3,3'}), 125.67 ppm (C^{8,8'}), 108.17 ppm (C^{7,7'}), 35.79 ppm (C^{11,11'}), 30.75 ppm (C^{12,12'}). Anal. Calcd for C₃₀H₃₆CrK₂O₃; C: 68.92%; H: 6.94%, Found; C: 68.53%; H 7.43%.

Synthesis of [^tBuOCO]Cr(THF)₃ (9). In a nitrogen filled glove box (368 mg (1.04 mmol) of CrCl₂Me(THF)₃ was dissolved in 20 mL of THF. In a separate vial 542 mg (1.04 mmol) of **10** was dissolved in 20 mL of THF. The solution of **10** was then added dropwise to the CrCl₂Me(THF)₃ solution with stirring at room temperature and stirred for 5 h. All volatiles were removed in vacuo. Toluene was added and the solution was filtered. The filtrate was evaporated to dryness to provide an oil that was dissolved in a minimal amount of THF and cooled to -35 °C to yield 292 mg **9** as a green crystalline solid (44% yield). ¹H NMR (300 MHz, benzene-*d*₆, δ): 8.23 ppm (br s), 4.90 ppm (br s), 1.45 ppm (br s), -7.48 ppm (br s), -13.30 ppm (br s). Selected IR data of **9** (neat film): ν (cm⁻¹) 1390 (s), 1250 (s), 1260 (w), 1125 (w), 1063 (m), 1010 (m), 850 (s), 840 (w), 812 (w). Anal. Calcd for C₃₈H₄₈CrO₅; C: 71.67%; H: 7.60%, Found; C: 71.24%; H 8.16%.

Synthesis of {[^tBuOCO]Cr=O(THF) (8). In a nitrogen filled glove box, 89 mg (0.140 mmol) of **3** was dissolved in 15 mL of toluene. The reaction vessel was fitted with a y-adaptor and attached to a Schlenck line. The solution was degassed and then O₂ gas was admitted (1 atm). The solution quickly turned purple then over the course of 2 h turned red brown. The solution was then degassed and the volatiles removed in vacuo yielding 58 mg of **8** as brown powder (41% yield). The solid can be recrystallized by

dissolving the brown powder in a minimal amount of toluene and cooling the solution to -35 °C. ^1H NMR (300 MHz, benzene- d_6 , δ) 11.4 ppm (br s), 9.04 ppm (br s), 4.23 ppm (br s), 2.37 ppm (br s), 1.21 (br s). Selected IR data of **8** (neat film): ν (cm^{-1}) 1577 (w), 1549 (w), 1471 (w), 1410 (s), 1359 (w), 1320 (w), 1242 (m), 1193 (m), 1110 (w), 1054 (w), 988 (s), 875 (m), 858 (w), 838 (w). Anal. Calcd for $\text{C}_{60}\text{H}_{70}\text{Cr}_2\text{O}_8 \cdot \text{C}_7\text{H}_8$; C: 72.15%; H: 7.05%, Found; C: 72.37%; H 6.68%.

As shown in Figure 2, upon addition of **1** to a dispersion of potassium hydride in THF, hydrogen gas evolves and the solution forms a pale yellow color. Filtration of any excess KH followed by concentration of the filtrate provides a colorless oil. Repeated trituration of the oil with pentane removes residual THF and provides **10** as a white solid in 75% yield. The salt is unstable in chloroform and sparingly soluble in benzene. ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR spectra are instead obtained in THF- d_8 . Resonances corresponding to phenol-OH's are absent in the ^1H NMR spectrum of **10**. The protons from the THF associated with **9** appear as broad singlets upfield from the residual protio solvent peaks at 3.59 and 1.74 ppm. The *ipso*-proton on the central aromatic ring resonates at 7.98 ppm and includes a weak $^4J_{\text{H-H}}$ coupling to the two *meta*-protons. A ^1H - ^1H COSY spectrum also reveals this coupling in addition to a long-range coupling to the *para*-proton. COSY, HETCOR, and gHMBC spectroscopic techniques enable identification assignment of each proton and carbon resonance.

A single-crystal X-ray experiment elucidates the exact orientation of the K^+ ion, the THF molecule, and the terphenyl fragment. Figure 5 depicts the molecular structure of **8**, Table 1 lists selected bond lengths and angles, and Table 2 contains pertinent crystallographic data. The asymmetric unit consists of one dipotassium unit and a THF molecule. This unit propagates in a two-dimensional sheet via K^+ -arene π -interactions. Three close contacts between K1 and the central aromatic ring ($d(\text{K1-C1}) = 2.966(3)$, $d(\text{K1-C2}) = 3.128(3)$, and $d(\text{K1-C6}) = 3.162(3)$ Å) and three additional connections to an adjacent molecule in the lattice ($d(\text{K1-C7}) = 3.282(3)$, $d(\text{K1-C8}) = 2.895(3)$, and $d(\text{K1-C9}) = 3.218(3)$ Å) complete the two-dimensional network.

Table 1. Selected bond lengths and angles for compounds 10.

Bond Length (Å)		Angle (°)	
K1-C1	2.966(3)	O1-K1-O2	86.26(7)
K1-C2	3.128(3)	O1-K2-O2	88.61(7)
K1-C6	3.162(3)	K1-O1-K2	86.99(7)
K1-O1	2.780(2)	K1-O2-K2	92.50(7)
K1-O2	2.601(2)	K1-C1-H1	104.0(19)
K2-O1	2.673(3)	C2-C1-C6	122.2(3)
K2-O2	2.596(2)	O1-K1-C1	62.39(8)
K1-K2	3.7543(11)	O2-K1-C1	65.21(8)
K2-H1	2.93(3)	C2-K1-C6	45.86(8)

Table 2. X-ray crystallographic structure parameters and refinement data for 10.

empirical formula	C ₃₀ H ₃₆ K ₂ O ₃
formula weight	522.79
crystal system	Monoclinic
space group	C2/c
crystal dimensions (mm)	0.17 × 0.14 × 0.13
<i>a</i> (Å)	18.3698(16)
<i>b</i> (Å)	18.2304(15)
<i>c</i> (Å)	16.6555(14)
β (deg)	94.661(2)
volume (Å ³)	5559.3(8)
<i>Z</i> (Å)	8
absorption coeff (mm ⁻¹)	0.369
<i>F</i> (000)	2224
<i>D</i> _{calcd} (g/cm ³)	1.249
γ (Mo K α) (Å)	0.71073
Temperature (K)	173(2)
θ range (deg)	1.58 to 27.50
completeness to $\theta_{\max}$	99.7%
index ranges	-23 ≤ <i>h</i> ≤ 23, -23 ≤ <i>k</i> ≤ 14, -20 ≤ <i>l</i> ≤ 21
reflections collected	18576
indep reflections [<i>R</i> _{int}]	6361 [0.0591]
data/restraints/param	6361/61/328
final <i>R</i> ₁ indices [<i>I</i> > 2 σ (<i>I</i>)]	<i>R</i> ₁ = 0.0737, <i>wR</i> ₂ = 2203 [4418]
<i>R</i> indices (all data)	<i>R</i> ₁ = 0.0977, <i>wR</i> ₂ = 0.2376
largest diff peak/hole e.Å ⁻³	1.569/-0.662
goodness of fit on <i>F</i> ²	1.053

Synthesis and Characterization of [tBuOCO]Cr^{III}(THF)₃ (9). As shown in Figure 3 Upon addition of **10** to MeCrCl₂(THF)₃ species in THF, the solution instantly darkens from lime green to dark green. After 5 hours, solvent removal provides a green solid that is extracted into toluene and KCl removed by filtration. After removing the toluene in vacuo, the solids remaining are dissolved in minimal THF and cooled to -35 °C to induce crystallization and produces analytically pure **9** in 44% yield.

A ¹H NMR spectrum of **9** presented in Figure 6 reveals paramagnetically shifted and broadened resonances. The signals appear at 8.20 (*v*_{1/2} = 518), 1.45 (*v*_{1/2} = 1035; tBu), -7.44 (*v*_{1/2} = 645), -13.23 (*v*_{1/2} = 750), and -22.71 (*v*_{1/2} = 2520) ppm. No resonances were observed in the ¹³C{¹H} NMR spectrum of **9**. Unfortunately, the ¹H spectrum is not

useful to confirm the identity of **9**. However, enough information is available (location of resonances and $\nu_{1/2}$) to determine if subsequent reactions lead to new chromium-containing products. The resonance at 1.45 ppm corresponds to the combined protons from the ^tBu groups and the three THF ligands bound to Cr(III).

1,2-dimethoxyethane yields crystals that retain solvent for a time sufficiently long enough to obtain X-ray data. As a consequence of recrystallizing in DME, two THF molecules are replaced in the Cr(III) coordination sphere with one DME. The structure of **9-DME** is presented in Figure 7, Table 3 list pertinent bond lengths and angles, and Table 4 provides the crystallographic data.

Table 3. Selected bond lengths and angles for compounds 9–DME.

Bond Length (Å)		Angle (°)	
Cr1–C1	1.994(3)	O1–Cr1–O2	177.42(9)
Cr1–O1	1.927(2)	C1–Cr–O5	93.67(11)
Cr1–O2	1.909(2)	C1–Cr1–O3	98.75(11)
Cr1–O3	2.086(2)	C1–Cr1–O4	174.98(11)
Cr1–O4	2.188(2)	O3–Cr1–O4	76.46(9)
Cr1–O5	2.049(2)	O1–Cr1–C1–C2	29

Table 4. X-ray crystallographic structure parameters and refinement data for 9-DME.

empirical formula	C ₃₈ H ₅₅ CrO ₇
formula weight	675.82
crystal system	Monoclinic
space group	P2(1)/n
crystal dimensions (mm)	0.19 × 0.14 × 0.03
<i>a</i> (Å)	10.9230(8)
<i>b</i> (Å)	24.4200(19)
<i>c</i> (Å)	13.5436(11)
β (deg)	97.571(2)
volume (Å ³)	3581.1(5)
<i>Z</i> (Å)	4
absorption coeff (mm ⁻¹)	0.366
<i>F</i> (000)	1452
D _{calcd} (g/cm ³)	1.253
γ (Mo K α) (Å)	0.71073
Temperature (K)	173(2)
θ range (deg)	1.67 to 25.00
completeness to $\theta_{\max}$	99.9%
index ranges	-12 ≤ <i>h</i> ≤ 12, -27 ≤ <i>k</i> ≤ 28, -11 ≤ <i>l</i> ≤ 16
reflections collected	20013
indep reflections [<i>R</i> _{int}]	6290 [0.0726]
data/restraints/param	6290/0/415
final <i>R</i> ₁ indices [<i>I</i> > 2 σ (<i>I</i>)]	<i>R</i> ₁ = 0.00564, <i>wR</i> ₂ = 0.1271 [4466]
<i>R</i> indices (all data)	<i>R</i> ₁ = 0.0888, <i>wR</i> ₂ = 0.1417
largest diff peak/hole e.Å ⁻³	0.555/-0.308
goodness of fit on <i>F</i> ²	1.038

5 The complex consists of a distorted octahedral Cr(III) coordinated by the OCO³⁻ pincer, THF, and DME ligands. The OCO pincer ligand adopts a pseudo C₂-symmetric orientation. As a consequence, the THF and DME ligands are canted towards open space as evidenced by O–Cr–O angles greater than 90° ($\angle$ O1–Cr1–O5 = 93.67(11)° and $\angle$ O2–Cr1–O3 = 92.12(9)°). A strong trans influence from the Cr–C1 bond (*d*(Cr1–C1) = 1.909(2) Å) causes a 0.1 Å elongation in the Cr1–O4 bond length (*d*(Cr1–O4) = 2.188(2) Å) compared to Cr1–O3 (*d*(Cr1–O3) = 2.086(2) Å). As expected, shorter bonds form between the Cr(III) ion and the alkoxide attachments (*d*(Cr1–O1) = 1.927(2) Å and *d*(Cr1–O2) = 1.909(2) Å).

10

Synthesis of $[\text{tBuOCO}]\text{Cr}^{\text{V}}=\text{O}(\text{THF})$ (8). As shown in Figure 4, when **9** is treated with an excess of O_2 (1 atm) in toluene the Cr(V) $[\text{tBuOCO}]\text{Cr}^{\text{V}}=\text{O}(\text{THF})$ (**8**) forms. As the O_2 is delivered, the solution color turns from bright green to purple and then to red-brown. After removal of all volatiles a brown powder remains, which can be further purified if necessary by recrystallization from a minimal amount of toluene at -35°C . Though the ^1H NMR spectrum of the brown powder does not reveal its identity, the signals indicate the presence of a new complex. The spectrum of the mixture in C_6D_6 exhibits several broad paramagnetically shifted peaks at 11.4, 9.0, 4.3, 2.4, and 1.2 ppm. Addition of THF to the NMR tube causes the signal at 1.2 ppm to grow, and is attributed to a bound THF that can exchange with free THF. Figure 8 shows the IR spectrum of **8**. The $\text{Cr}^{\text{V}}-\text{O}$ stretch appears as a strong absorption at 988 cm^{-1} .

Synthesis of **2, R = Mesityl).** In a glovebox, as shown in Figure 9, 1,3-benzene-bis(aniline) (0.95 g, 3.7 mmol), tris(dibenzylideneacetone)-dipalladium (0) (0.168 g, 0.183 mmol), *rac*-2,2'-bis(diphenylphosphino)-1,1'-binaphthyl (0.284 g, 0.455 mmol), sodium tert-butoxide (1.077 g, 11.2 mmol), 40 mL of toluene, and mesityl bromide (6.84 mL, 44 mmol) were combined in a 250 mL, 2-neck round-bottom flask. The reaction mixture was refluxed under argon for 2 days with completion of the reaction confirmed by TLC (10 % ethyl acetate: 90 % hexanes). The reaction was quenched with 10 mL of water, the solvents removed under vacuum and dried for 2 hours. The crude product was extracted with pentanes and filtered. Colorless crystalline solid was obtained from cool pentanes overnight (730 mg, 38 %). ^1H NMR (300 MHz, C_6D_6) δ (ppm): 7.99 (s, 1H, Ar-H), 7.59 (d, $J = 7.6\text{ Hz}$, 2H, Ar-H), 7.35 (d, $J = 7.6\text{ Hz}$, 1H, Ar-H), 7.27-7.33 (m, 2H, Ar-H), 7.03 (t, $J = 7.6\text{ Hz}$, 2H, Ar-H), 6.81 (d, $J = 6.9\text{ Hz}$, Ar-H), 6.76 (s, 4H, Ar-H), 6.41 (d, $J = 7.6\text{ Hz}$, 2H, Ar-H), 5.29 (bs, 2H, N-H), 2.15 (s, 6H, p-Mes- CH_3), 2.03 (s, 12H, o-Mes- CH_3) as shown in Figure 10. $^{13}\text{C}\{^1\text{H}\}$ NMR (75.36 Hz, C_6D_6) δ (ppm): 144.4 (s, C, Ar), 141.7 (s, C, Ar), 136.7 (s, C, Ar), 136.4 (s, C, Ar), 135.9 (s, C, Ar), 131.3 (s, C, Ar), 131.2 (s, C, Ar), 130.3 (s, C, Ar), 129.9 (s, C, Ar), 129.0 (s, C, Ar), 128.0 (s, C, Ar), 118.8 (s, C, Ar), 112.4 (s, C, Ar), 21.4 (s, p-Mes- CH_3), 18.7 (s, o-Mes- CH_3) as shown in Figure 11. DIP-CIMS, $[\text{M}-\text{H}]^+ = 497.2948$ ($\text{C}_{36}\text{H}_{38}\text{N}_2$, theoretical = 497.2957).

Synthesis of $N,N'-(2,2'-(1,3\text{-phenylene})\text{bis(ethane-2,1-diyl)})\text{bis}(3,5\text{-bis(trifluoromethyl)aniline})$ (3** – R = 3,5-bis(trifluoromethylphenyl).** As shown in Figure 12, to a 100 mL round bottom flask charged with toluene (50 mL) were added 1,3-bis-

(2-aminoethyl)benzene (1.50 g, 9.15 mmol), 3,5-bis(trifluoromethyl)bromobenzene (5.37 g, 2 equiv, 18.3 mmol), Pd₂(dba)₃ (0.13 g, 0.5%, 0.142 mmol), *rac*-BINAP (0.23 g, 1.5%, 0.357 mmol), and NaO^tBu (2.64 g, 27.5 mmol). After refluxing for 72 h under argon the solution was filtered hot through celite and volatiles were removed *in vacuo*. Nonvolatile products were dissolved in hot pentanes and filtered through celite. The final product was recrystallized twice from pentane at -20 °C. Yield 2.1 g (3.57 mmol, 39%). ¹H NMR (300 MHz, C₆D₆, δ): 7.22 (s, 2H, *Ar-H*), 7.11 (t, *J* = 7.64 Hz, 1H, *Ar-H*), 6.81 (dd, *J* = 7.64, 1.70 Hz, 2H, *Ar-H*), 6.72 (s, 1H, *Ar-H*), 6.48 (s, 4H, *Ar-H*), 3.13 (t, *J* = 5.52 Hz, 2H, *NH*), 2.73 (dt, *J* = 6.94 Hz, 4H, -NHCH₂CH₂-), 2.38 (t, *J* = 6.94 Hz, 4H, -NHCH₂CH₂Ar) as shown in Figure 13. ¹³C{¹H} NMR (128.39 Hz, C₆D₆, δ): 35.5 (s, -NHCH₂CH₂Ar), 44.6 (s, NHCH₂CH₂Ar), 110.4 (s, aromatic), 112.3 (s, aromatic), 122.9 (s, aromatic), 126.5 (s, aromatic), 127.6 (s, aromatic), 129.8 (s, aromatic), 133.0 (q, *J*_{C-F} = 32.7 Hz, -CF₃), 139.8 (s, ArCCH₂-), 149.2 (s, ArCNH) as shown in Figure 14. HRMS calculated (found) for C₂₆H₂₀F₁₂N₂ (M+H)⁺: 589.1508 (589.1537).

X-ray Data for (3 – R = 3,5-bis(trifluoromethylphenyl)) were collected at 173 K on a Siemens SMART PLATFORM equipped with A CCD area detector and a graphite monochromator utilizing MoK_α radiation (λ = 0.71073 Å). Cell parameters were refined using up to 8192 reflections. A full sphere of data (1850 frames) was collected using the ω-scan method (0.3° frame width). The first 50 frames were re-measured at the end of data collection to monitor instrument and crystal stability (maximum correction on I was < 1 %). Absorption corrections by integration were applied based on measured indexed crystal faces.

The structure was solved by the Direct Methods in *SHELXTL6*, and refined using full-matrix least squares. The non-H atoms were treated anisotropically, whereas the hydrogen atoms were calculated in ideal positions and were riding on their respective carbon atoms. The asymmetric unit consists of two chemically equivalent but crystallographically independent molecules. They differ by the orientations of the side aryl rings with respect to the central one. Out of the eight CF₃ groups, six of them are disordered and were refined in two parts each. A total of 890 parameters were refined in the final cycle of refinement using 10886 reflections with *I* > 2σ(*I*) to yield *R*₁ and *wR*₂ of 6.33% and 13.87%, respectively. Refinement was done using *F*². The Ortep drawing is given in Figure 15

Table 5. Crystal data, structure solution and refinement for (3 – R = 3,5-bis(trifluoromethylphenyl)).

identification code	ac04
empirical formula	C ₂₆ H ₂₀ F ₁₂ N ₂
formula weight	588.44
<i>T</i> (K)	173(2)
λ (Å)	0.71073
Crystal system	Triclinic
space group	P-1
<i>a</i> (Å)	11.9159(14)
<i>b</i> (Å)	14.003(2)
<i>c</i> (Å)	16.6156(18)
α (deg)	75.343(2)
β (deg)	71.727(2)
γ (deg)	74.212(2)
<i>V</i> (Å ³)	2490.3(5)
<i>Z</i>	4
ρ_{calcd} (Mg mm ⁻³)	1.570
Crystal size (mm ³)	0.11 x 0.08 x 0.05
abs coeff (mm ⁻¹)	0.156
<i>F</i> (000)	1192
θ range for data collection	1.31 to 22.50
limiting indices	$-11 \leq h \leq 12, -15 \leq k \leq 13, -12 \leq l \leq 17$
no. of reflns colld	10886
no. of ind reflns (<i>R</i> _{int})	6488 (0.1286)
completeness to $\theta = 27.49^\circ$	99.8 %
absorption corr	Integration
refinement method	Full-matrix least-squares on <i>F</i> ²
data / restraints / parameters	6488 / 0 / 890
<i>R</i> ₁ , ^a <i>wR</i> ₂ ^b [<i>I</i> > 2 σ]	0.0633, 0.1387
<i>R</i> ₁ , ^a <i>wR</i> ₂ ^b (all data)	0.1350, 0.1701
GOF ^c on <i>F</i> ²	0.896
largest diff. peak and hole	0.310 and -0.263 e.Å ⁻³

^a $R_1 = \sum ||F_o| - |F_c|| / \sum |F_o|$. ^b $wR_2 = (\sum (w(F_o^2 - F_c^2)^2) / \sum (w(F_o^2)^2))^{1/2}$. ^c $GOF = (\sum w(F_o^2 - F_c^2)^2 / (n - p))^{1/2}$ where *n* is the number of data and *p* is the number of parameters refined.

Synthesis of (4 – R = 3,5-bis(trifluoromethyl)phenyl). As shown in Figure 16, to an anhydrous DMF (15 mL) solution of bis(bromomethylene)anthracene (2.81 g, 7.72 mmol) and K₂CO₃ was added 3,5-bis(trifluoromethyl)aniline (2.65 ml, 17.0 mmol). The mixture was stirred for 15 h and then water was added to form a yellow oil layer. The yellow oil was extracted with CHCl₃, washed with water, brine, dried with MgSO₄, and then filtered. After removing all volatiles, an oil forms with some yellow solid. The oil was dissolved in hexanes and the yellow solid (bis(bromomethylene)anthracene starting material) was removed by filtration. The oil contained 4 – R = 3,5-bis(trifluoromethyl)phenyl, aniline, and DMF which was separated by flash column chromatography (100 g SiO₂, 1:1 CHCl₃:pentane) to provide pure 4 – R = 3,5-bis(trifluoromethyl)phenyl as a pale yellow solid. Yield 2.36 g (3.57 mmol, 46%). ¹H NMR (300 MHz, C₆D₆, δ): 8.31 (s, 1H, Ar H), 8.23 (s, 1H, Ar H), 7.77 (d, *J* = 6 Hz, 2H, Ar H), 7.0-7.4 (m, 6H, Ar H), 6.49 (s, 4H, *o*-CHCF₃), 4.00 (d, *J* = 3.0 Hz, 4H, N-CH₂-), 3.23 (t, *J* = 3.0 Hz, 2H, NH) as shown in Figure 17. ¹³C{¹H} NMR (128.39 C₆D₆, δ): 46.8 (s, CH₂), 111.0 (s, aromatic), 112.1 (s, aromatic), 119.1 (s, aromatic), 125.8 (s, aromatic), 127.0 (s, aromatic), 129.1 (s, aromatic), 129.6 (s, aromatic), 130.5 (s, aromatic), 132.7 (s, aromatic), 133.1 (q, *J*_{CF} = 33.1 Hz, CF₃), 133.6 (s, aromatic), 149.1 (s, *i*-C-NH) as shown in Figure 18. HRMS calculated (found) for C₃₂H₂₀F₁₂N₂ (M⁺): 660.14 (660.1465).

Synthesis of (5 – R = Si^{*i*}Pr₃). As shown in Figure 19, a solution of 1,8-anthracenediamine, (208 mg, 1 mmol) in tetrahydrofuran (10 mL) was cooled to -32 °C, and ^{*n*}BuLi (0.81 mL, 2.5 M solution in hexane, 2.01 mmol) was then added. The reaction mixture was slowly warmed to room temperature and was stirred for 2 h. The resulting mixture was then again cooled to -32 °C, and to this ¹Pr₃SiCl (0.43 mL, 2.01 mmol) was added very slowly over a period of 5 minutes. The mixture was slowly brought to room temperature and was continuously stirred for 12 h. The solvent was removed under vacuum and dried 2 hours. The crude product was extracted into hexanes (10 mL). The solvent was removed under vacuum to yield a viscous dark brown oil (450 mg, 87 %). ¹H NMR (300 MHz, C₆D₆) δ (ppm): 8.60 (s, 1H, Ar-*H*), 8.29 (s, 1H, Ar-*H*), 7.50 (d, *J* = 8.6 Hz, 2H, Ar-*H*), 7.34 (t, *J* = 8.6 Hz, *J* = 7.3 Hz, 2H, Ar-*H*), 6.94 (d, *J* = 7.3 Hz, 2H, Ar-*H*), 4.12 (bs, 2H, N-*H*), 1.31-1.40 (m, 6H, ¹Pr-CH), 1.17 (d, *J* = 7.0 Hz, 27H, ¹Pr-CH₃), 1.02 (d, *J* = 3.0 Hz, 9H, ¹Pr-CH₃) as shown in Figure 20. ¹³C{¹H} NMR (75.36 Hz, C₆D₆) δ

(ppm): 143.6 (s, C, Ar), 133.9 (s, C, Ar), 128.8 (s, C, Ar), 126.7 (s, C, Ar), 126.0 (s, C, Ar), 119.7 (s, C, Ar), 110.7 (s, C, Ar), 19.3 (s, ¹Pr-CH), 18.2 (¹Pr-CH₃), 14.3 (s, ¹Pr-CH), 13.4 (¹Pr-CH₃) as shown in Figure 21. ESI-MS, [M]⁺ = 520.3685 (C₃₂H₅₂N₂Si₂, theoretical = 520.3664), [M+H]⁺ = 521.3754 (C₃₂H₅₃N₂Si₂, theoretical = 521.3742), [2M+H]⁺ = 1041.7443 (C₆₄H₁₀₅N₄Si₄, theoretical = 1041.7411).

Synthesis of 1,3-bis(7-methyl-1H-indol-2-yl)benzene (6). As shown in Figure 21, the synthesis involves a Madelung synthesis of indoles. A stirred solution of diamide (0.54g 1.45 mmol) in 100 mL of dry THF maintained under a N₂ atmosphere was maintained at an internal temperature of -35 °C and treated dropwise with 5.8 mmol of ⁿBuLi. The stirred mixture was kept at ambient temperature overnight, cooled in an ice bath, and treated dropwise with of 2 N HCl. The organic layer was separated and the aqueous layer washed with ether. The combined organic layers were dried with anhydrous MgSO₄, filtered, and concentrated in vacuo to obtain an off white solid.

Synthesis of Cr(III) NCN pincer ligand (11). An {NCNLi₂}₂ dimer can be prepared from butyl lithium and compound 7 – R = 3,5-bis-(*i*-propyl)phenyl. As shown in Figure 22, the {NCNLi₂}₂ salt (1.0 g, 1.1 mmols) was dissolved in diethyl ether and cooled to -80 °C. In a separate vial CrMeCl₂(THF)₃ (0.75 g, 2.1 mmol) was dissolved in diethyl ether and cooled to -80 °C. The Cr salt solution was slowly added to the NCNLi salt solution while maintaining the temperature at -80 °C. After the addition was complete, the combined solution warm to ambient temperature over a 30 minute period while stirring was maintained. The solution was filtered using a medium coarse fritted funnel and the filtrate was collected and volatiles were removed in vacuo to give a dark maroon solid. The solid was triturated with pentane, extracted into pentane and filtered through a fritted funnel. The filtrate was collected and the solvent removed in vacuo. The resulting dark maroon solid was redissolved in a minimal amount of pentane and cooled to -80 °C. A NCNCr(THF)₃ complex precipitate was isolated by filtration as a dark maroon microcrystalline powder. Yield 0.53 g (1.04 mmol, 49%). ¹H NMR (300 MHz, C₆D₆, δ): 26.05 (bs, *Ar-H*), 10.43 (bs, *Ar-H*), 4.17 (bs, THF), 1.23 (bs, overlapping THF and CH(CH₃)₂ protons), 0.82 (bs, *Ar-H*), -7.44 (bs, *Ar-H*) shown in Figure 23.

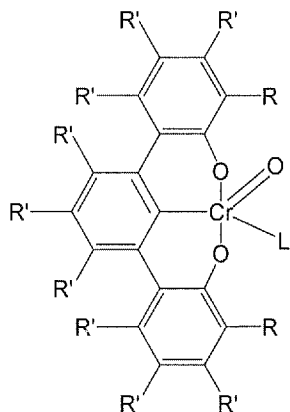
It should be understood that the examples and embodiments described herein are for illustrative purposes only and that various modifications or changes in light thereof

will be suggested to persons skilled in the art and are to be included within the spirit and purview of this application.

CLAIMS

We claim:

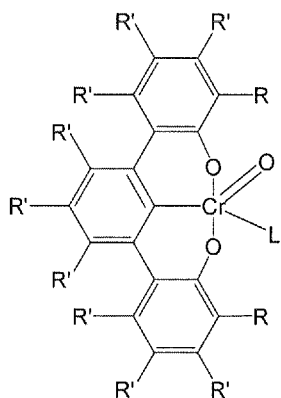
1. A Cr(V) OCO^{3-} trianionic pincer ligand complex comprising the structure:



where: R is independently a C3 to C12 alkyl or aryl group; R' is independently a H, C1 to C12 alkyl or aryl group; and L is an oxygen comprising ligand.

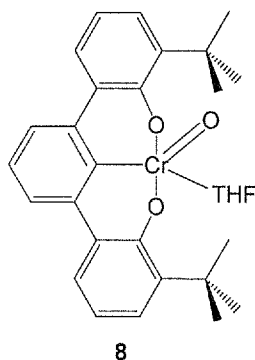
2. The pincer ligand complex of claim 1, wherein R is *t*-butyl, R' is H and L is tetrahydrofuran (THF).
3. A catalytic method for oxidation of a substrate comprising the steps of:
 - providing a Cr(V) OCO^{3-} trianionic pincer ligand complex;
 - providing a source of oxygen molecules; and
 - providing a substrate, wherein said substrate is oxidized and said Cr(V) OCO^{3-} trianionic pincer ligand complex is reduced by the substrate to give a Cr(III) complex and oxidizing said Cr(III) complex to said Cr(V) OCO^{3-} trianionic pincer ligand complex.
4. The method of claim 3, wherein said Cr(V) OCO^{3-} trianionic pincer ligand complex is provided in a solvent.
5. The method of claim 4, wherein said solvent comprises an aliphatic hydrocarbon, an aromatic hydrocarbon, or an ether.

6. The method of claim 3, wherein said source of oxygen comprises a gas comprising oxygen and an inert gas.
7. The method of claim 6, wherein said oxygen comprises 1 to 100 % of said source.
8. The method of claim 3, wherein said source of oxygen comprises air.
9. The method of claim, 3, wherein said substrate comprises a phosphine, amine, sulfide, alkene, alkane or a second metal complex.
10. The method of claim 3, wherein said substrate is triphenyl phosphine.
11. The method of claim 3, wherein said $\text{Cr(V)} \text{OCO}^{3-}$ trianionic pincer ligand complex comprises:



where: R is independently a C3 to C12 alkyl or aryl group; R' is independently a H, C1 to C12 alkyl or aryl group; and L is an oxygen comprising ligand.

12. The method of claim 3, wherein said $\text{Cr(V)} \text{OCO}^{3-}$ trianionic pincer ligand complex comprises:



13. A method of preparing a $\text{Cr(V)} \text{OCO}^{3-}$ trianionic pincer ligand complex comprising the steps of:

- providing an OCOH_3 pincer ligand;
- converting said OCOH_3 pincer ligand to an OCOHM_2 pincer complex;
- exchanging said OCOHM_2 pincer complex with a Cr(III) salt; and
- oxidizing said Cr(III) to Cr(V) .

14. The method of claim 13, wherein said OCOH_3 pincer ligand comprises $(^t\text{BuOCO})\text{H}_3$.

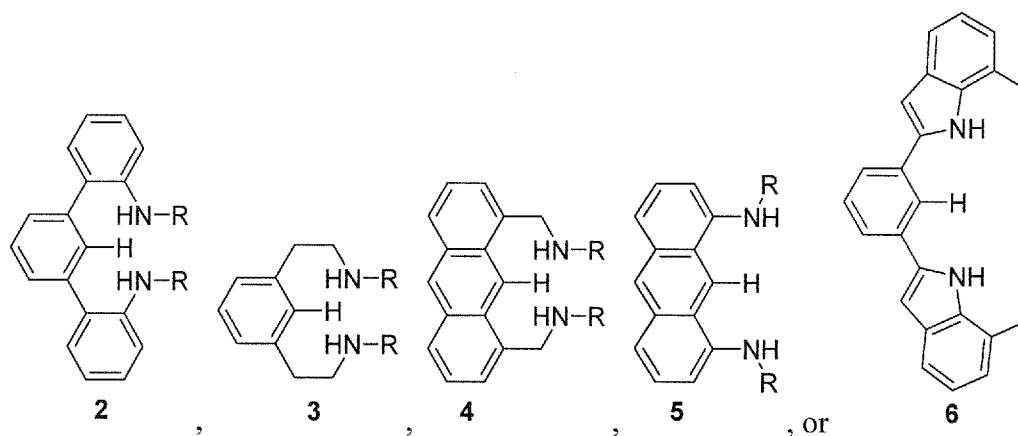
15. The method of claim 13, wherein said OCOHM_2 pincer complex comprises an OCOHK_2 pincer complex by reaction of said OCOH_3 pincer ligand with KH .

16. The method of claim 13, wherein said Cr(III) salt comprises $\text{CrCl}_2\text{Me(THF)}_3$.

17. The method of claim 13, wherein said oxidation comprises oxidation in an oxygen comprising gas.

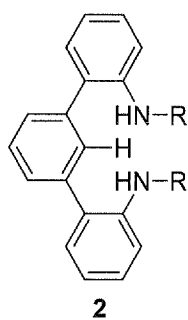
18. The method of claim 13, wherein said oxygen comprising gas is pure oxygen.

19. An NCN pincer ligand, comprising a structure selected from:



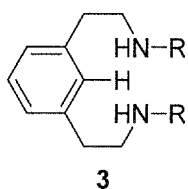
wherein R is 2,6-bis-(*i*-propyl)phenyl, 3,5-bis-(methyl)phenyl, 3,5-bis-(*i*-propyl)phenyl, 3,5-bis-(trifluoromethyl)phenyl, mesityl, or tri-*i*-propylsilyl.

20. The NCN pincer ligand of claim 19, wherein the ligand is:



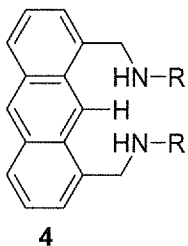
wherein R is mesityl.

21. The NCN pincer ligand of claim 19, wherein the ligand is:



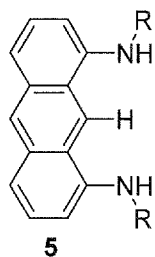
wherein R is 3,5-bis-(trifluoromethyl)phenyl.

22. The NCN pincer ligand of claim 19, wherein the ligand is:



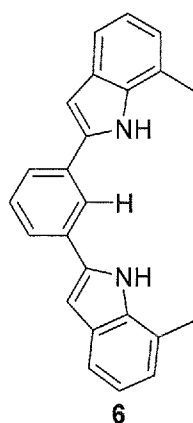
wherein R is 3,5-bis-(trifluoromethyl)phenyl.

23. The NCN pincer ligand of claim 19, wherein the ligand is:

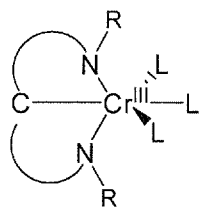


wherein R is tri-*i*-propylsilyl.

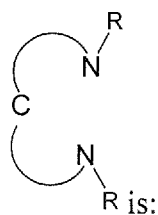
24. The NCN pincer ligand of claim 19, wherein the ligand is:

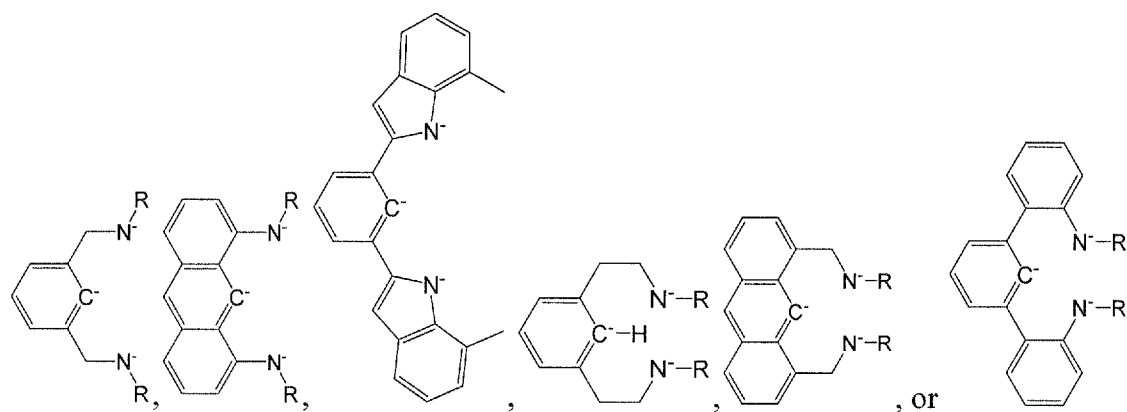


25. A Cr(III) NCN³⁻ trianionic pincer ligand complex comprising the structure:



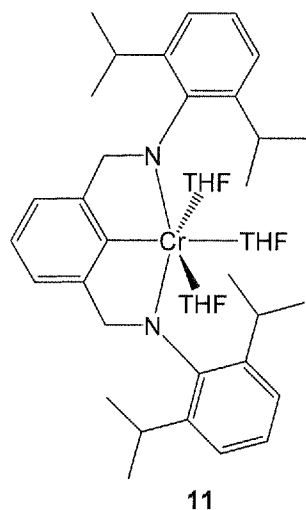
where:





where R is 2,6-bis-(*i*-propyl)phenyl, 3,5-bis-(methyl)phenyl, 3,5-bis-(trifluoromethyl)phenyl, 3,5-bis-(*i*-propyl)phenyl, mesityl, or tri-*i*-propylsilyl and L is any neutral coordinating ligand.

26. The Cr(III) NCN³⁻ trianionic pincer ligand complex of claim 25, wherein the complex is:



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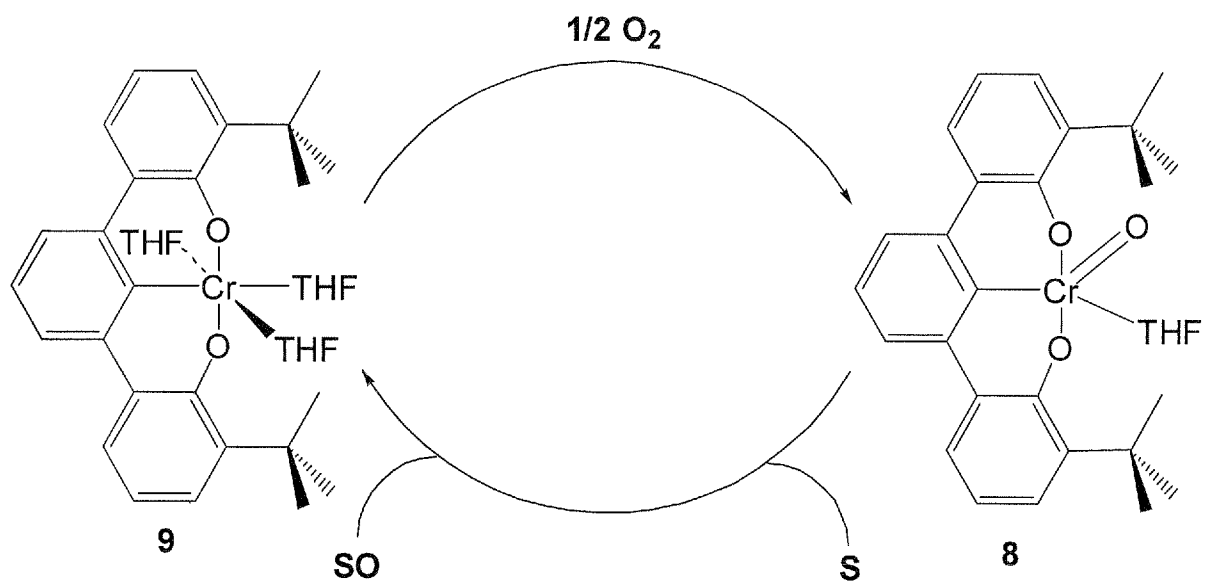


FIG. 1

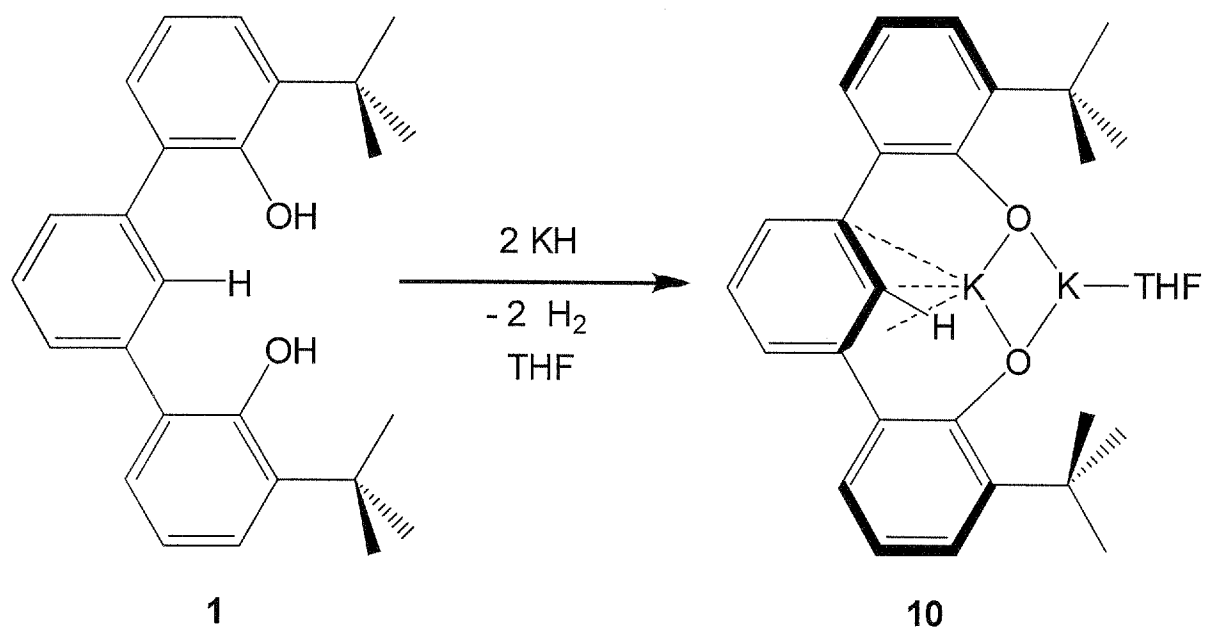


FIG. 2

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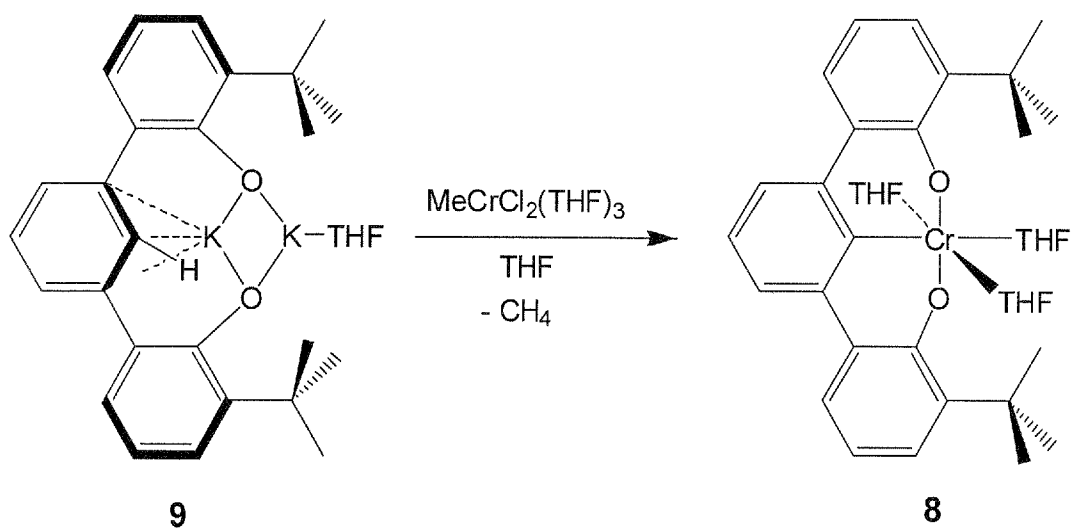


FIG. 3

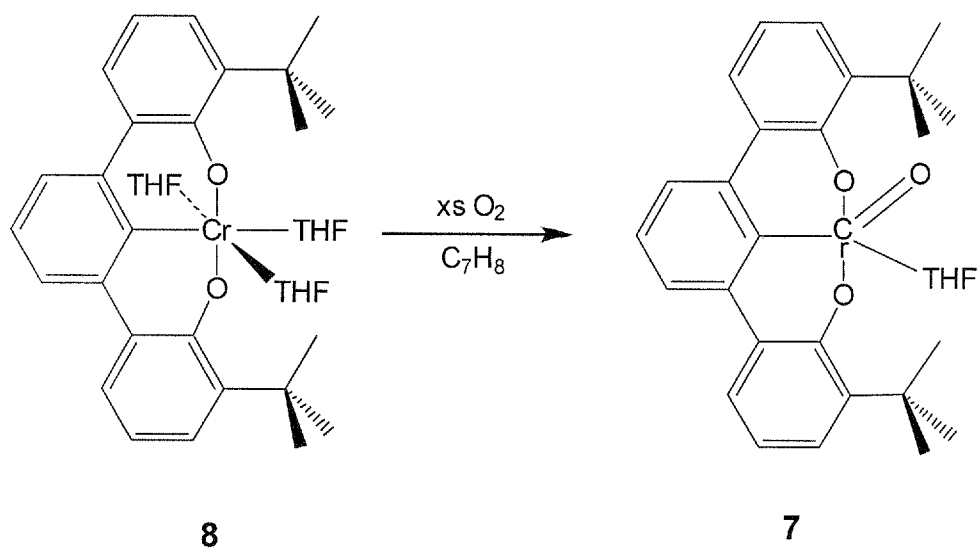
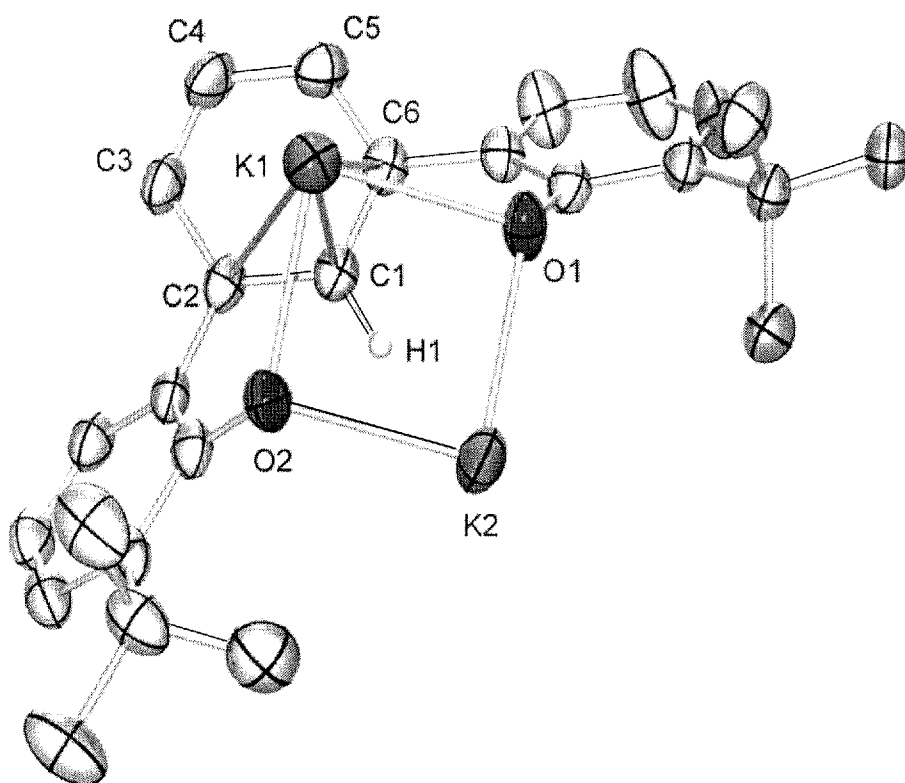
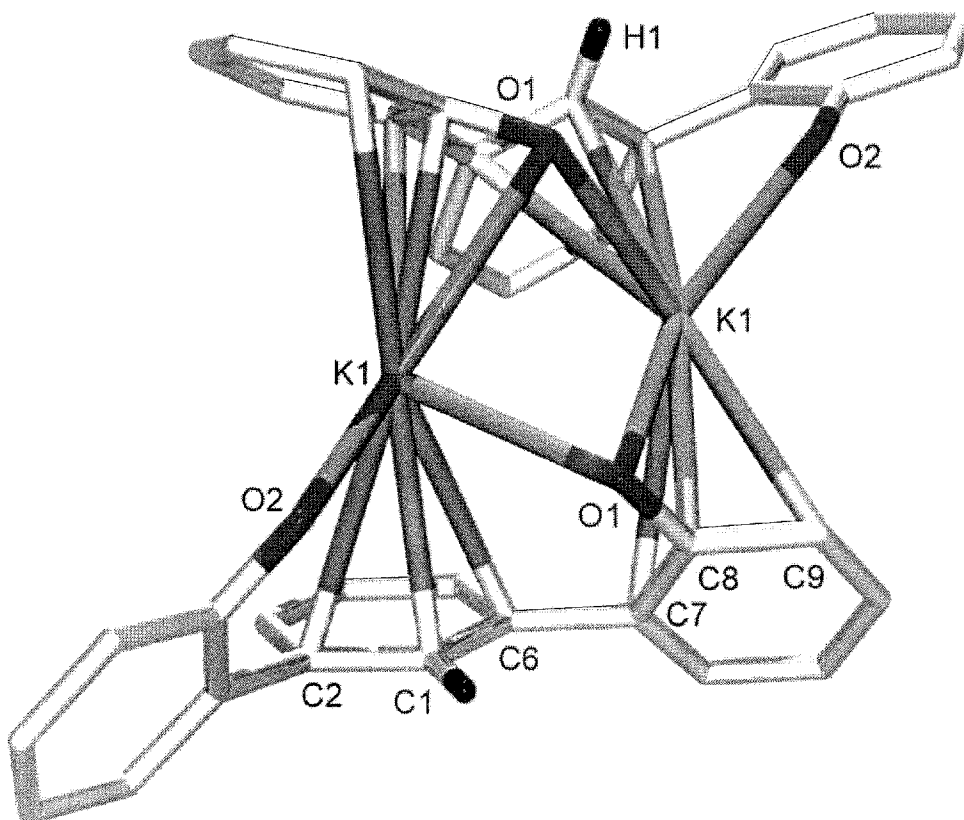


FIG. 4

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**FIG. 5A****FIG. 5B**

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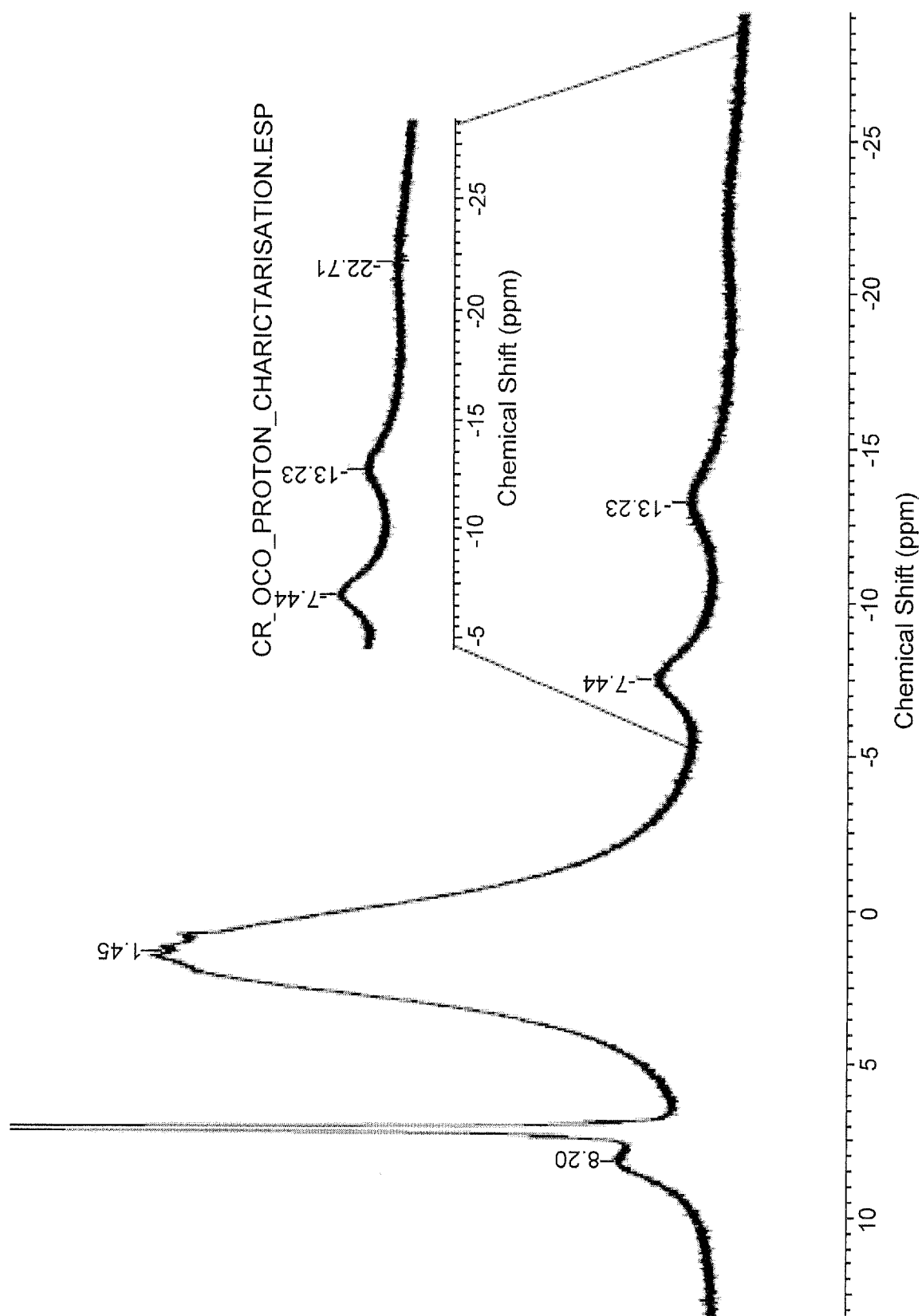


FIG. 6

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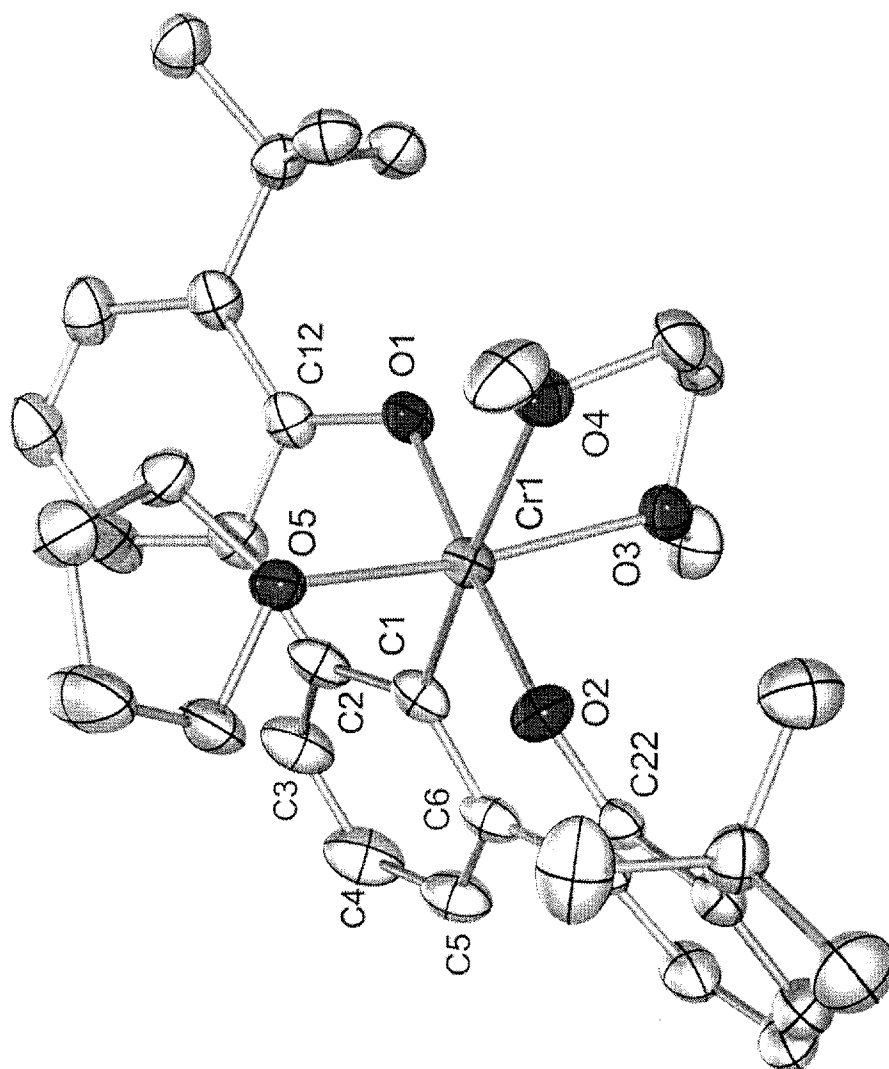


FIG. 7

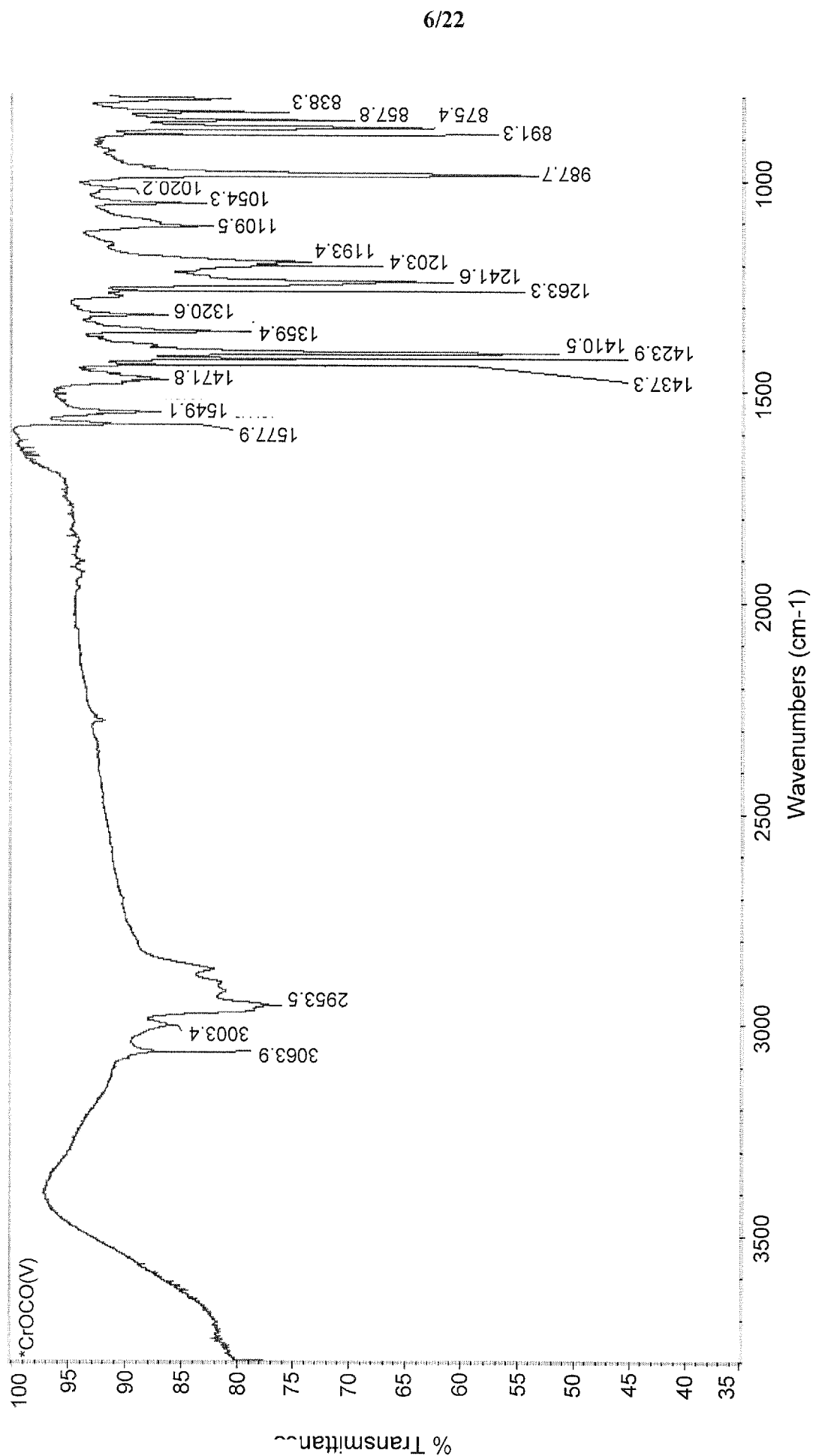


FIG. 8

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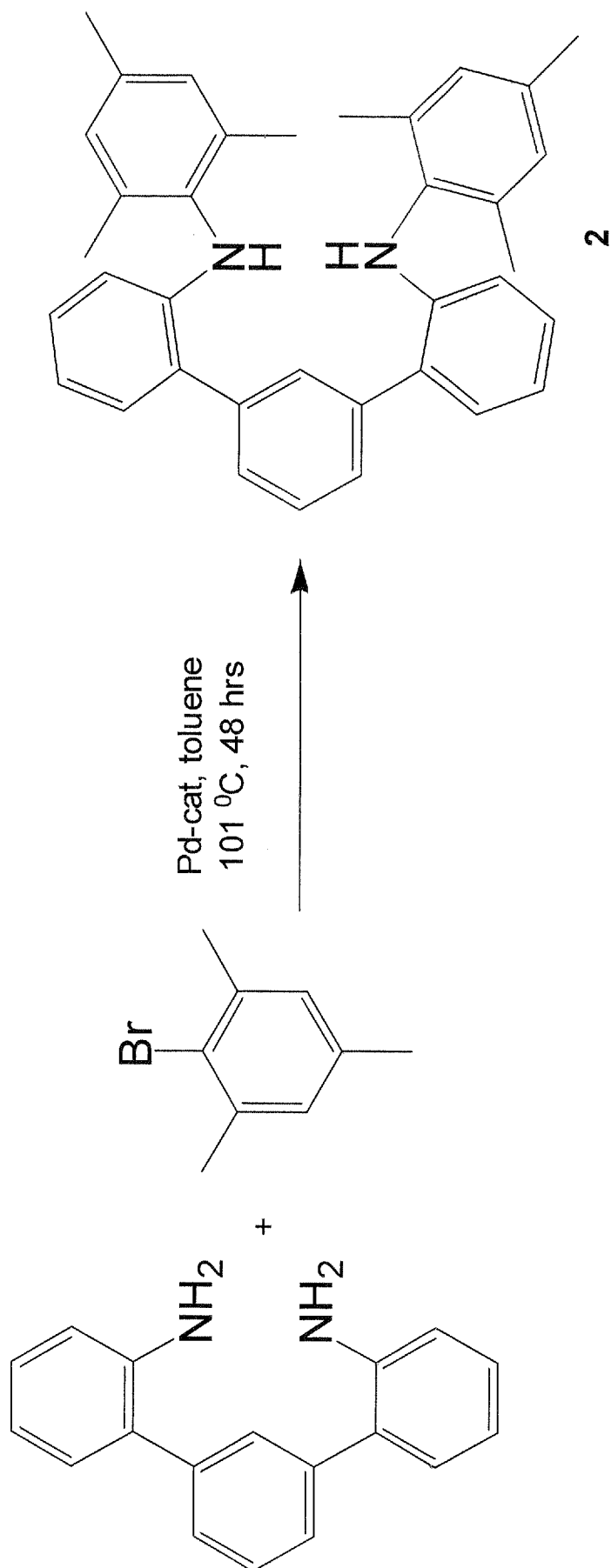
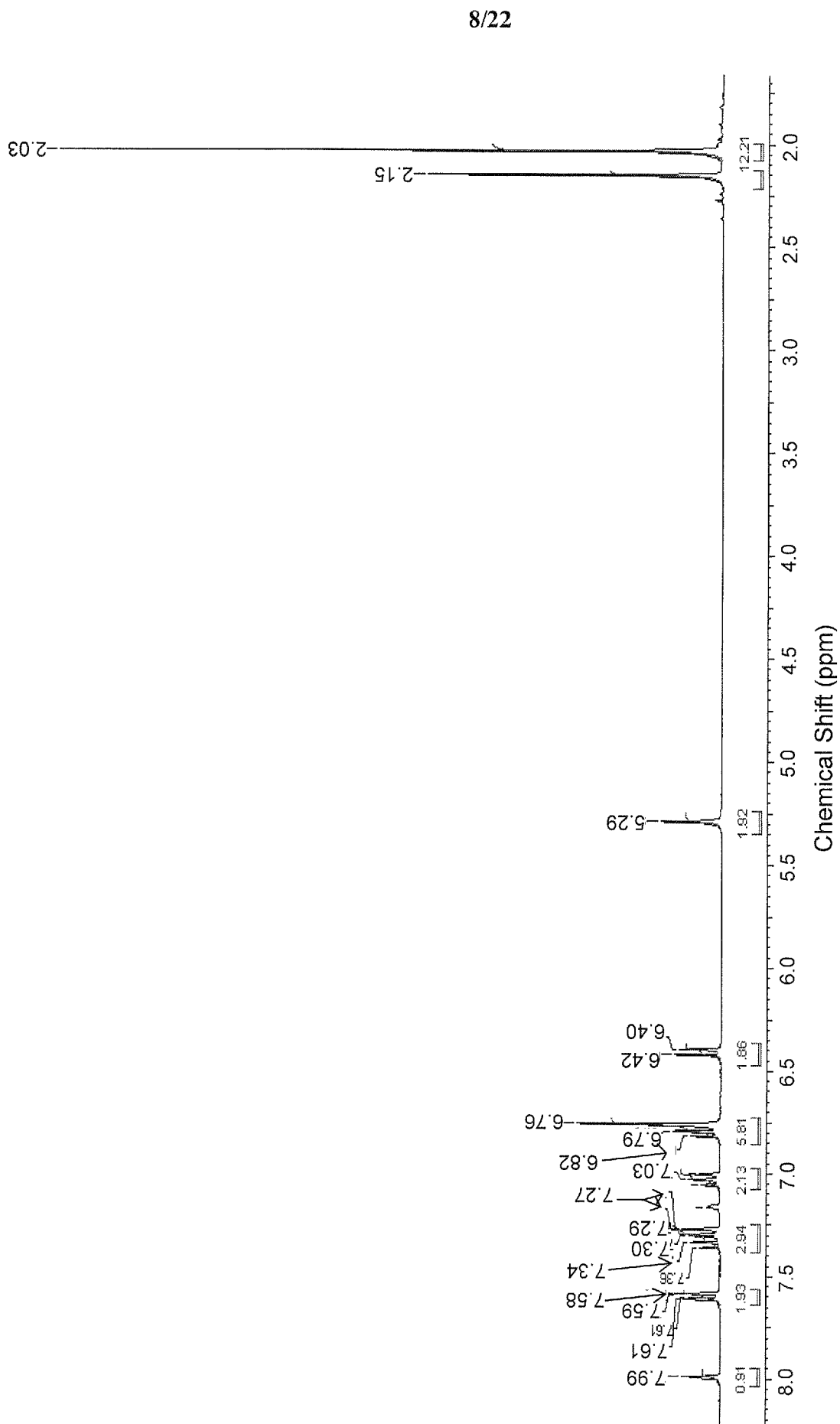


FIG. 9

**FIG. 10**

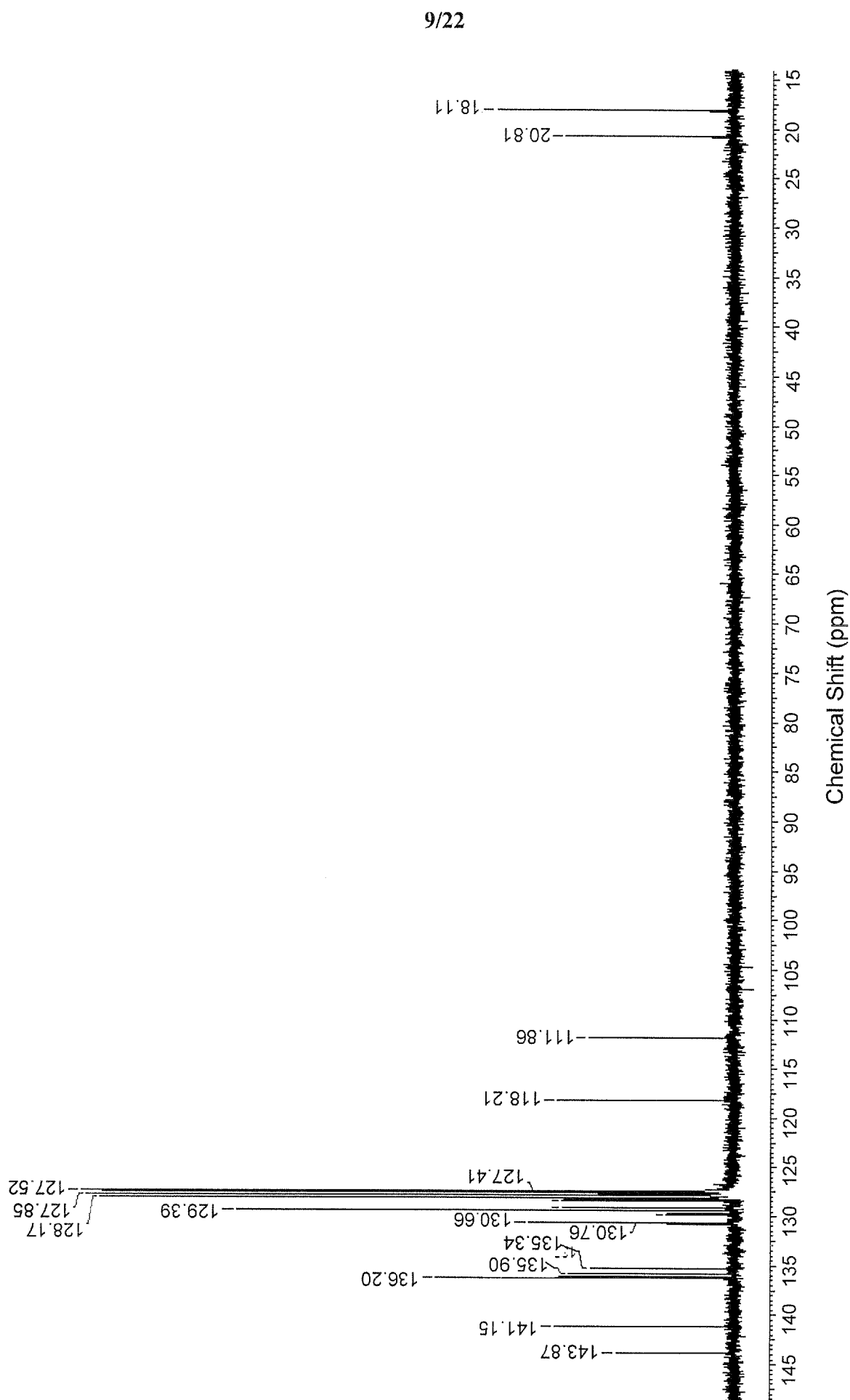


FIG. 11

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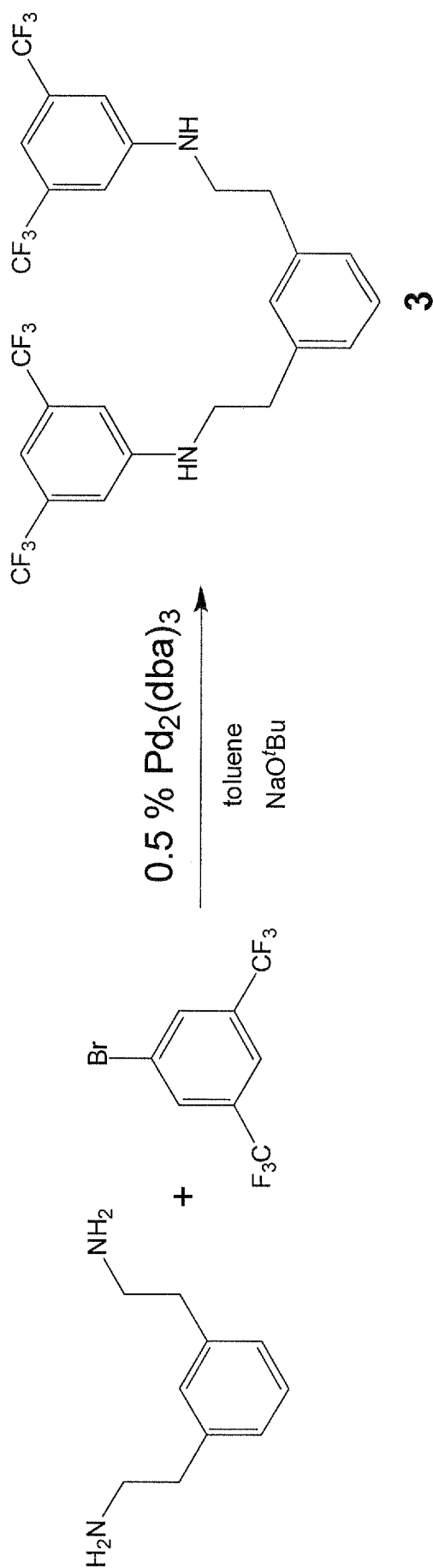
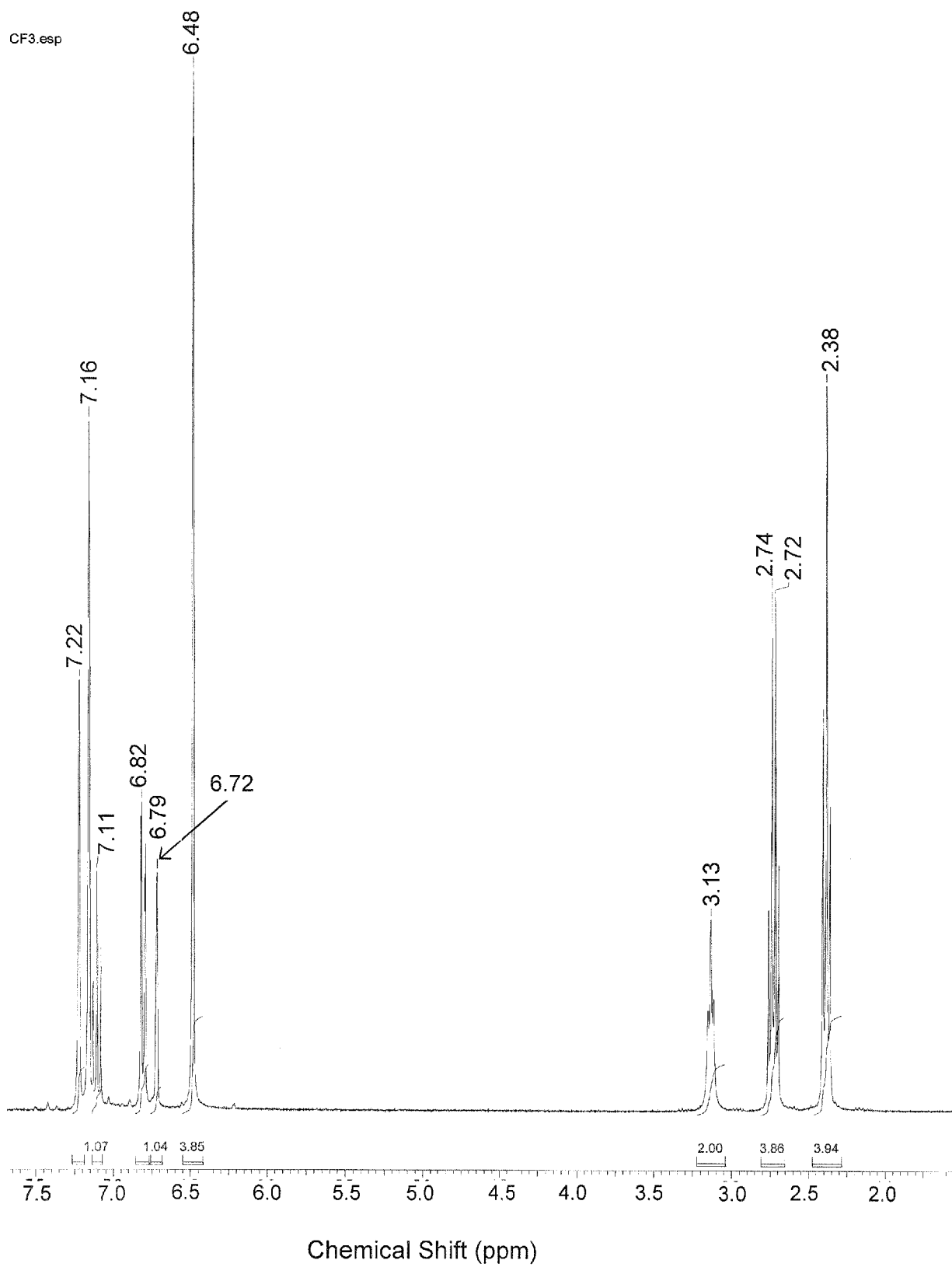
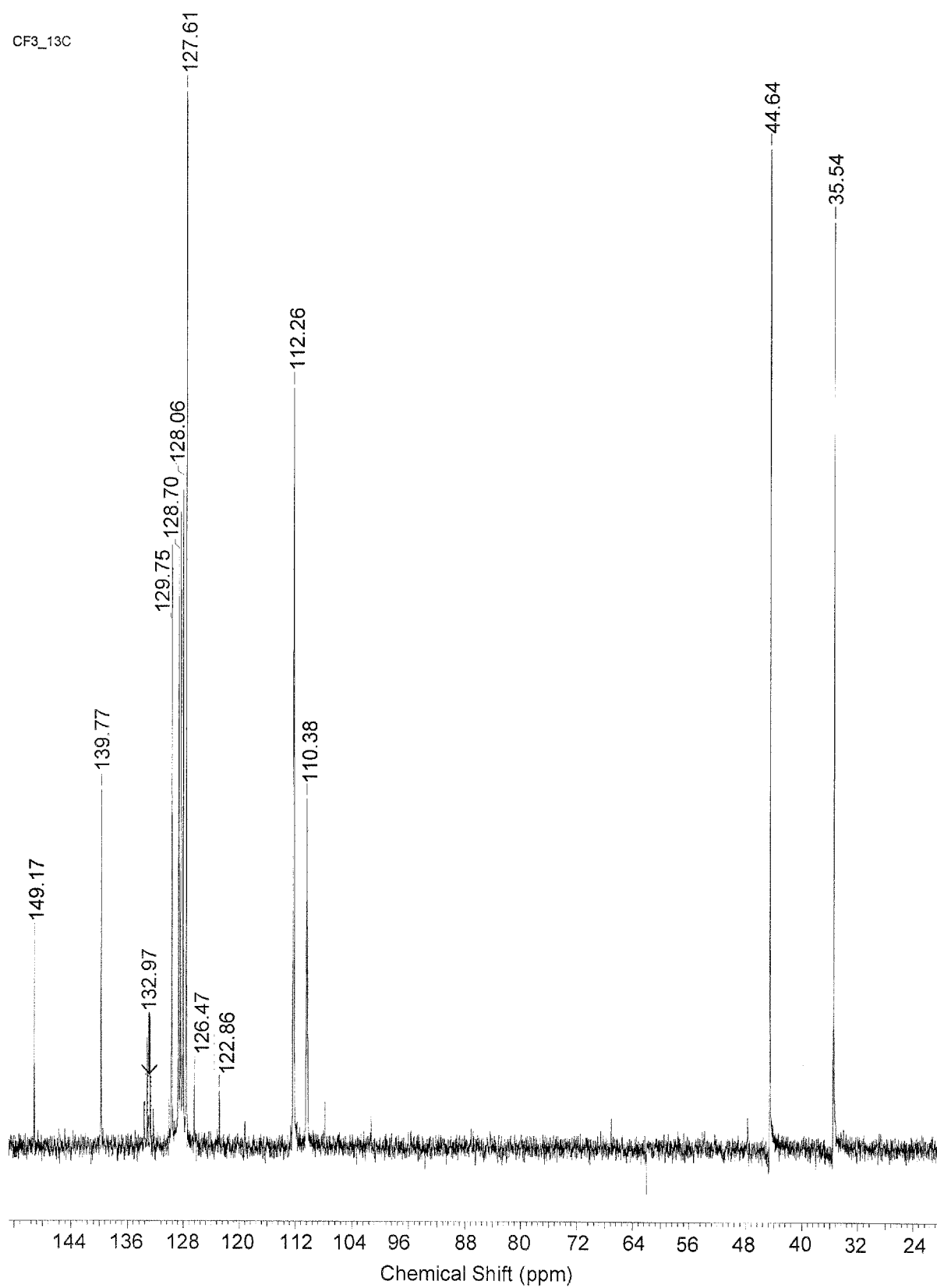


FIG. 12

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**FIG. 13**

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**FIG. 14**

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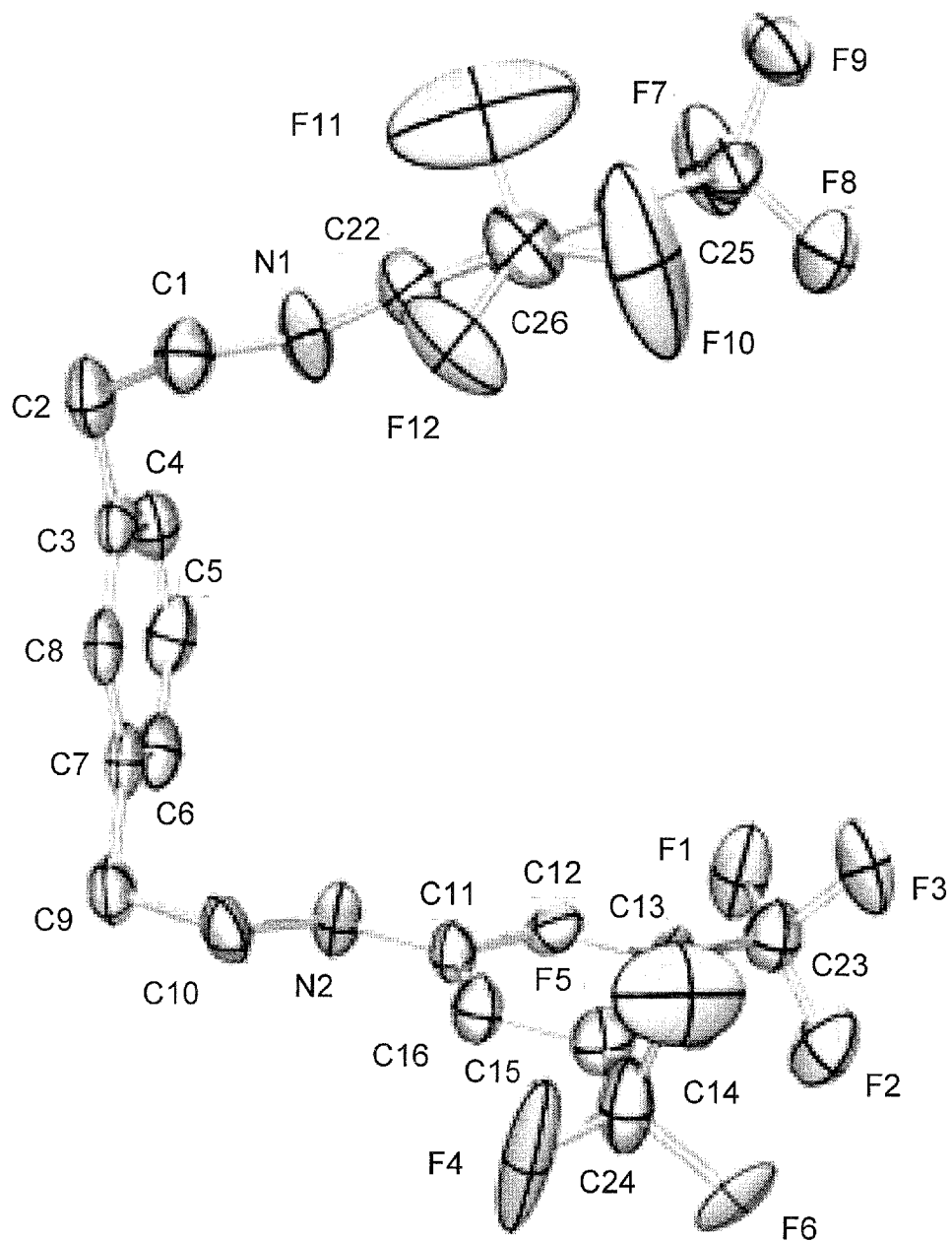
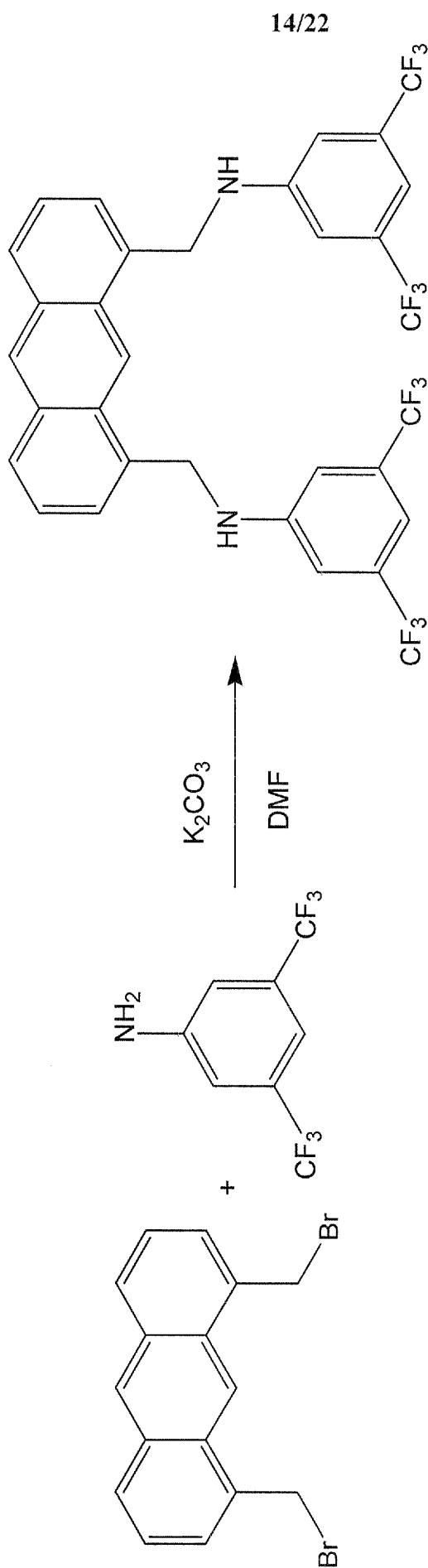


FIG. 15

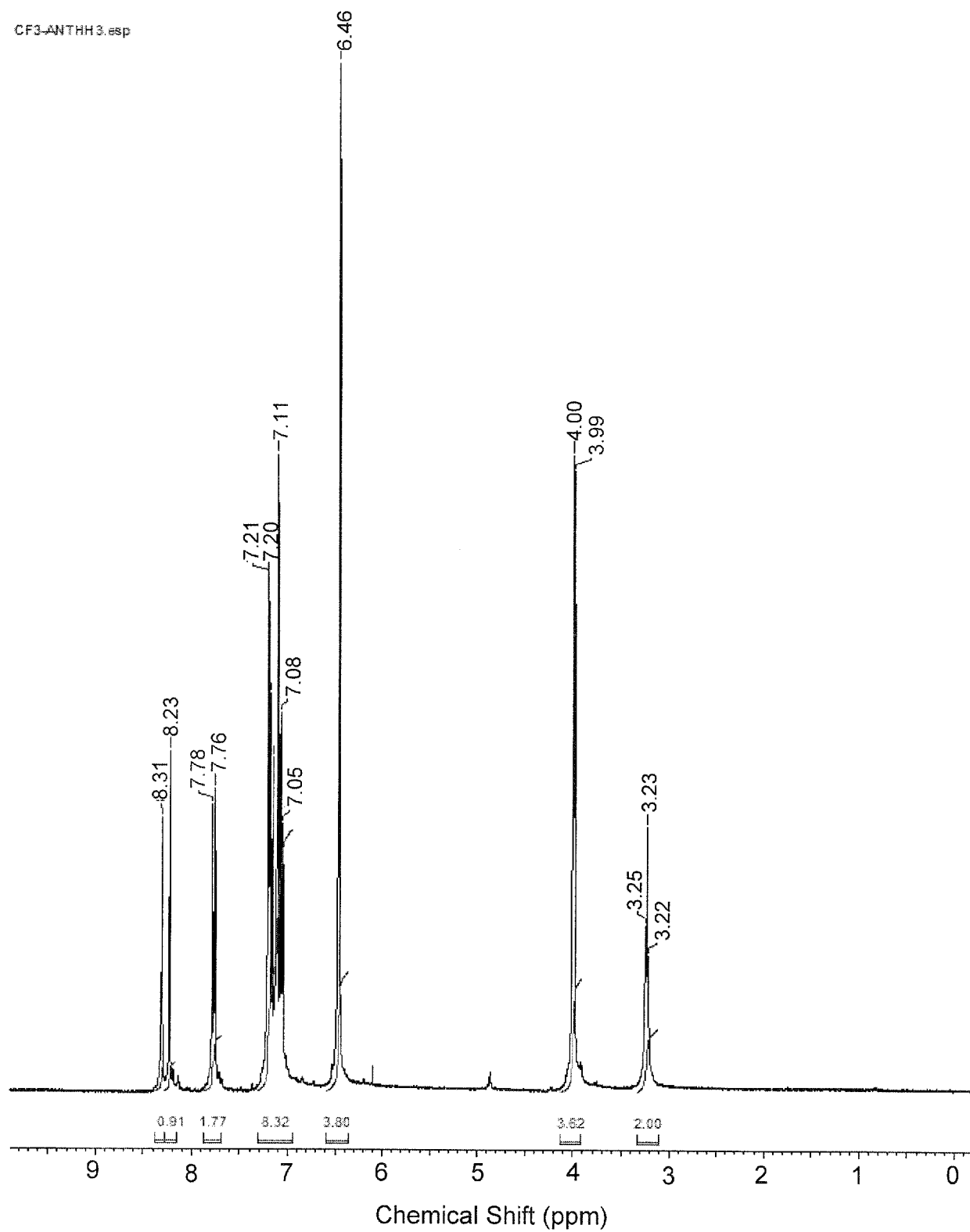


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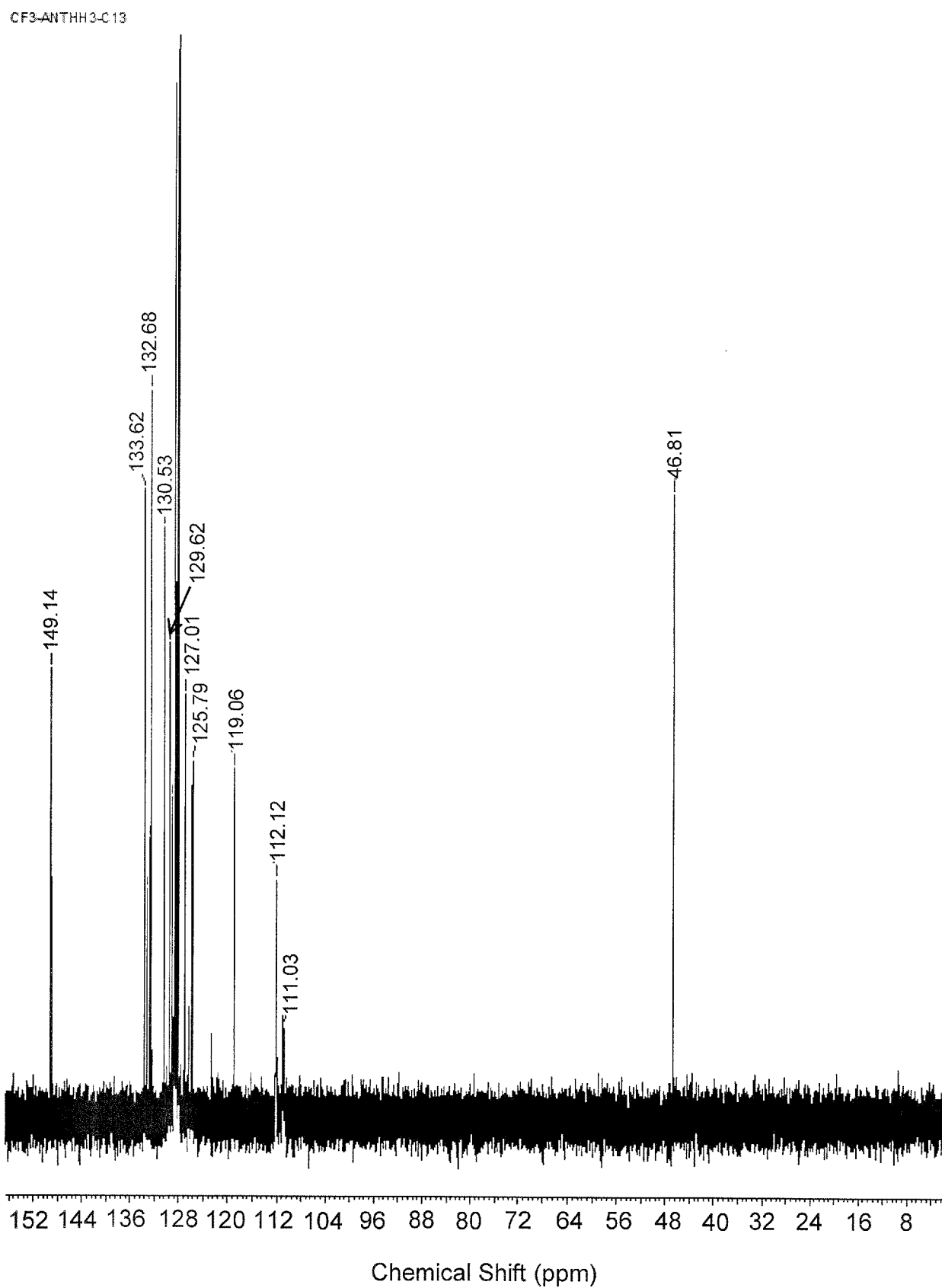
FIG. 16

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CF3-ANTHH 3.esp

**FIG. 17**

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**FIG. 18**

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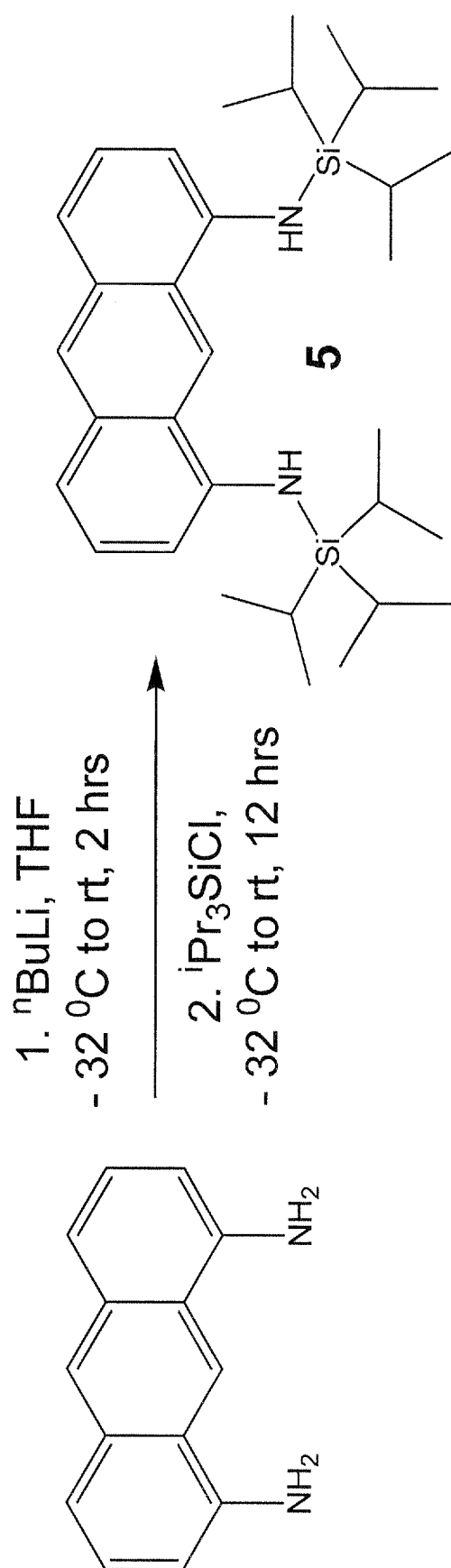
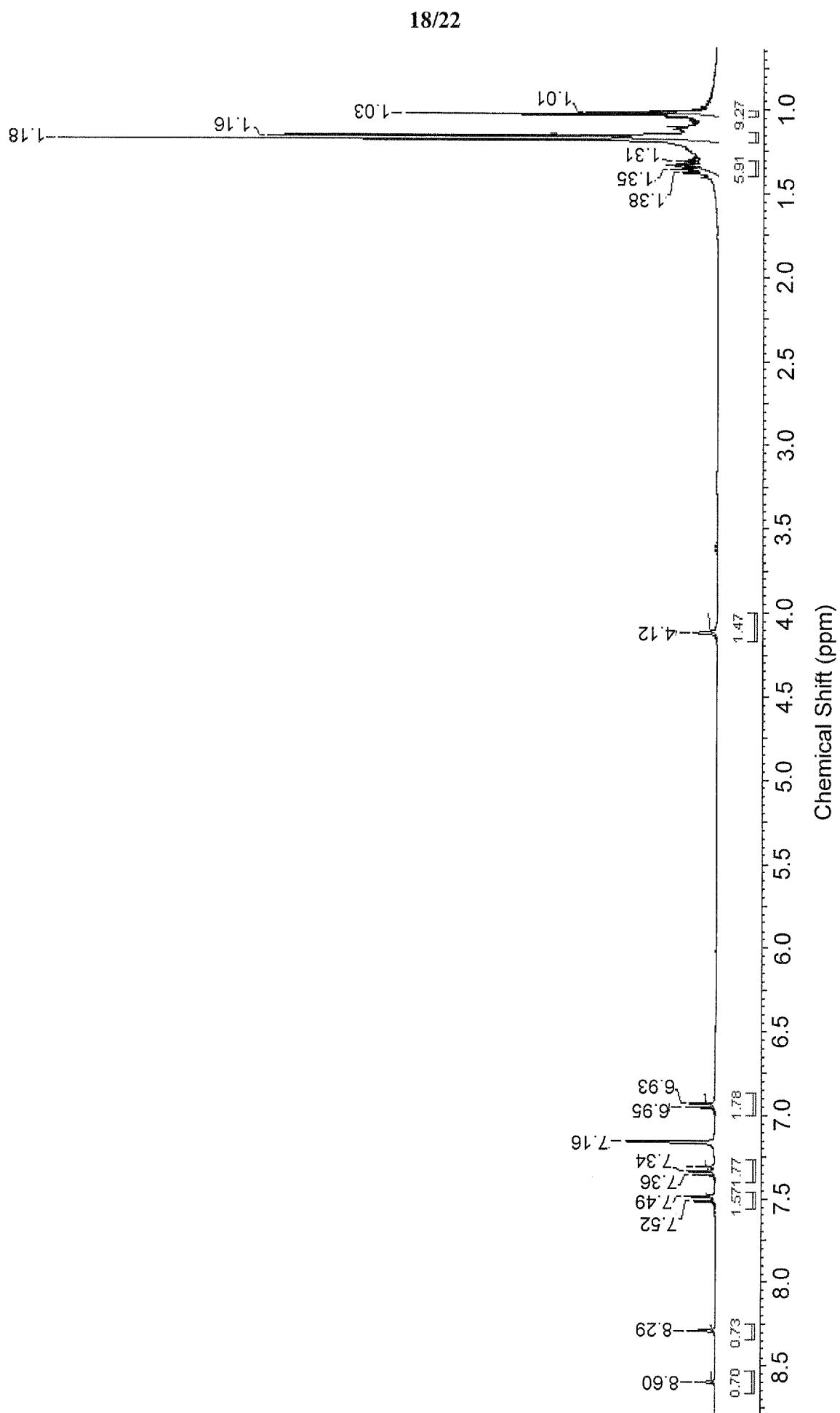


FIG. 19

**FIG. 20**

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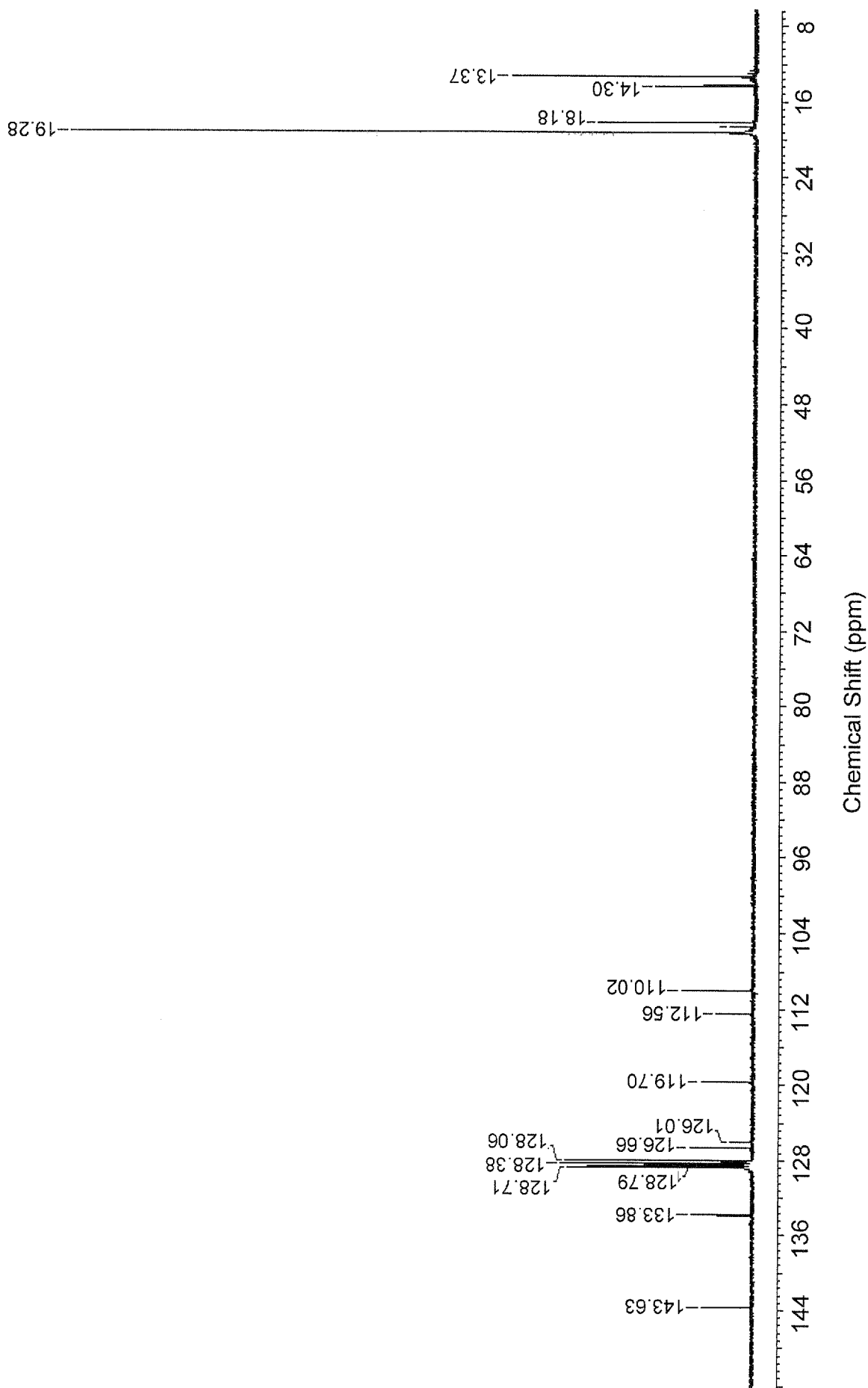


FIG.21

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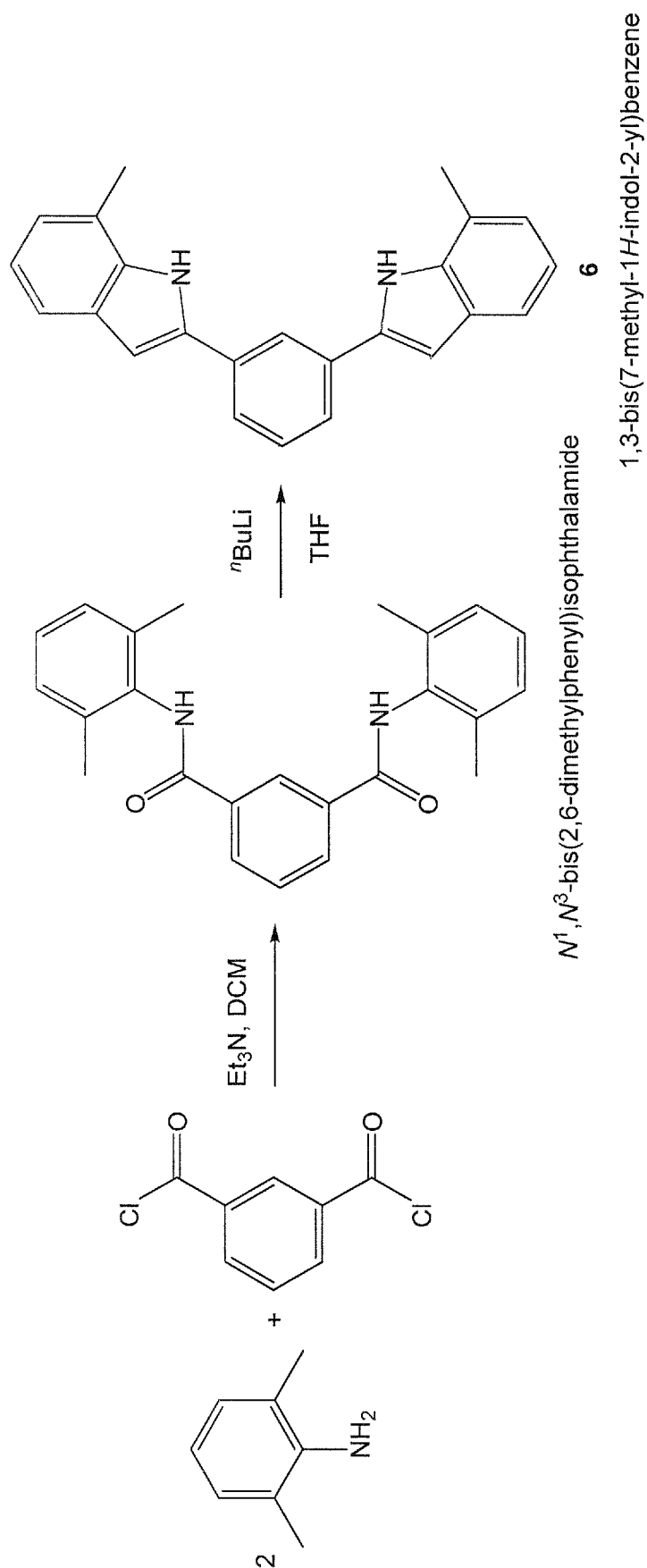


FIG.22

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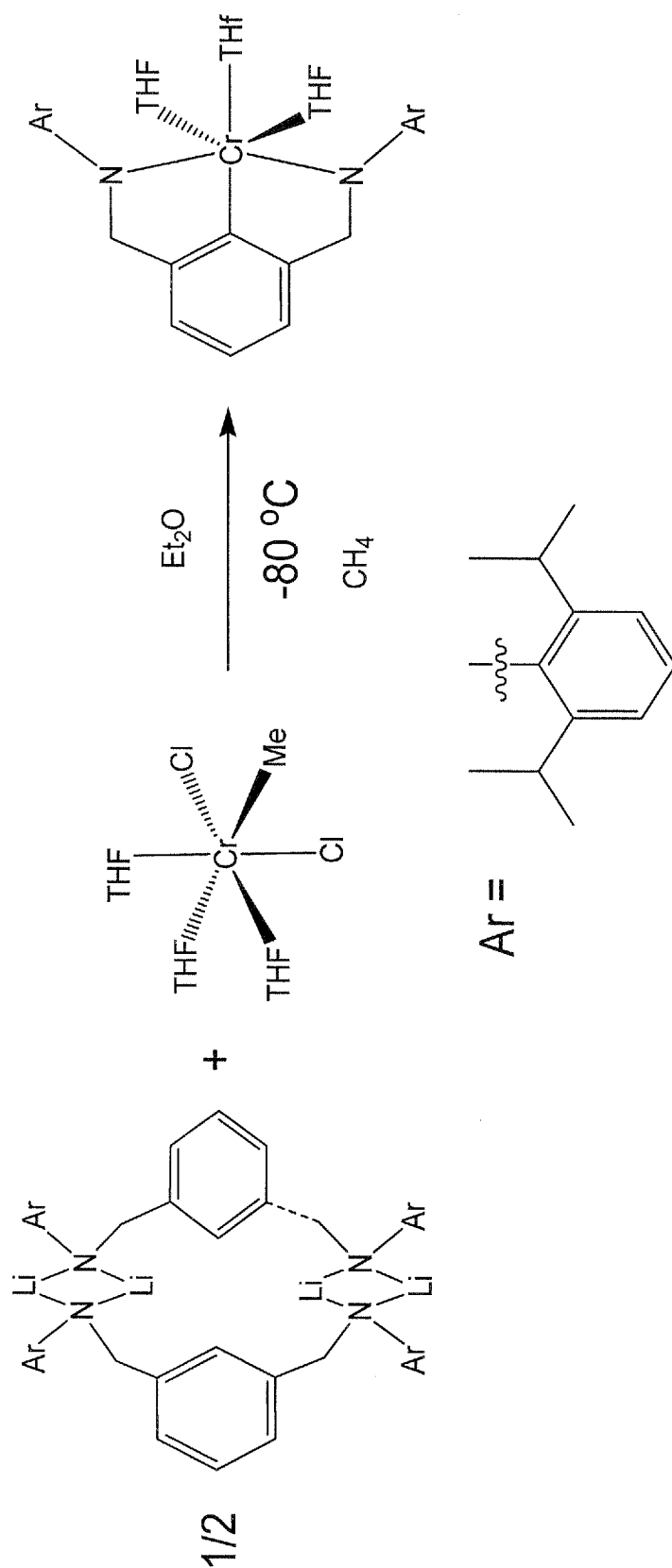


FIG.22

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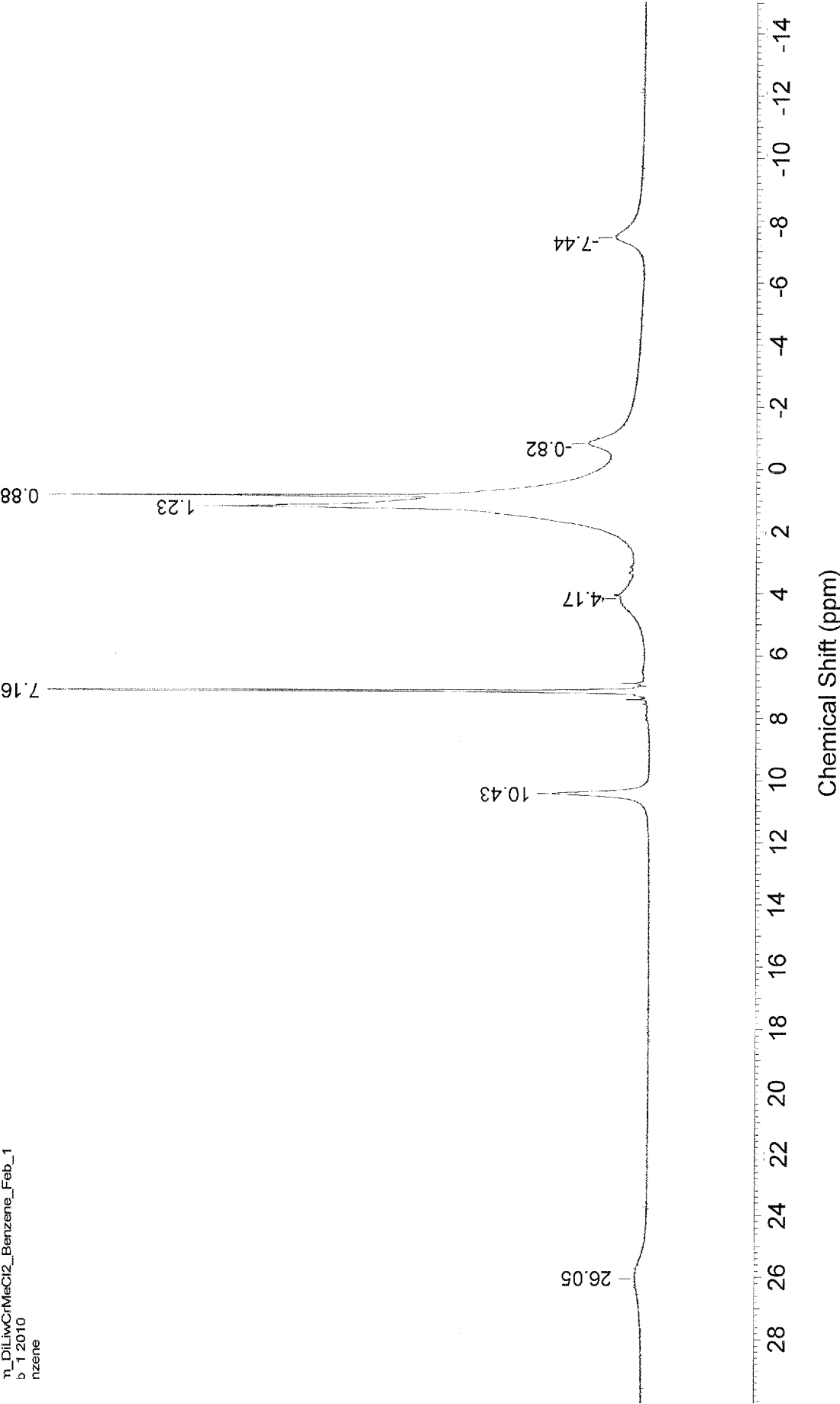


FIG. 23

n_DiIwCrMeCl2_Benzene_Feb_1
b 1 2010
nzene