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(54) **POLYMERIZATION CATALYSTS FOR PRODUCING POLYMERS WITH LOW LEVELS OF LONG CHAIN BRANCHING**

POLYMERISATIONSKATALYSATOREN ZUR HERSTELLUNG VON
LANGKETTENVERZWEIGUNGSARMEN POLYMEREN

CATALYSEURS DE POLYMERISATION POUR PRODUIRE DES POLYMERES A FAIBLE TENEUR
EN RAMIFICATION A LONGUE CHAINE

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(56) References cited:
EP-A- 0 853 086 US-A- 5 886 202
US-B1- 6 340 651

• **ALT, HELMUT G. ET AL: "C1-Bridged
fluorenylidene-indenylidene complexes of type
(C₁₃H₈-CR₂-C₉H₆-nR'_n)ZrCl₂ (n=0, 1; R=Me, Ph,
butenyl; R'=alkyl, alkenyl) as metallocene
catalysts for ethylene polymerization" JOURNAL
OF ORGANOMETALLIC CHEMISTRY , 562(2),
153-181 CODEN: JORCAI; ISSN: 0022-328X
JOURNAL OF ORGANOMETALLIC CHEMISTRY ,
562(2), 153-181 CODEN: JORCAI; ISSN:
0022-328X, 1998, XP004128016**
• **ALT, HELMUT G. ET AL: "ansa-Metallocene
complexes of type
(C₁₃H₈-SiR₂-C₉H₆-nR'_n)ZrCl₂ (n=0, 1; R=Me, Ph,
alkenyl; R'=alkyl, alkenyl): self-immobilized
catalysts for ethylene polymerization" JOURNAL
OF ORGANOMETALLIC CHEMISTRY , 562(2),
229-253 CODEN: JORCAI; ISSN: 0022-328X
JOURNAL OF ORGANOMETALLIC CHEMISTRY ,
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0022-328X, 1998, XP004128023**

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Description

BACKGROUND OF THE INVENTION

[0001] It is known that mono-1-olefins (α -olefins), including ethylene, can be polymerized with catalyst compositions employing titanium, zirconium, vanadium, chromium, or other metals, often combined with a solid oxide and in the presence of cocatalysts. These catalyst compositions can be useful for both homopolymerization of ethylene, as well as copolymerization of ethylene with comonomers such as propylene, 1-butene, 1-hexene, or other higher α -olefins. Therefore, there exists a constant search to develop new olefin polymerization catalysts, catalyst activation processes, and methods of making and using catalysts that will provide enhanced catalytic activities and polymeric materials tailored to specific end uses.

[0002] Polyethylene (PE) produced by any number of methods generally contains small to moderate amounts of long chain branched molecules. In some instances, long chain branching (LCB) is desired to improve bubble stability during film blowing or to enhance the processibility of resins prepared with metallocene catalysts. However for many uses, the presence of LCB is considered undesirable due to the increased elasticity that it typically imparts to the resins. Therefore the ability to control the LCB level in polyethylene using metallocene-based catalysts is a desirable goal.

[0003] One example of this need is seen in the use of bridged or *ansa*-metallocene catalysts, which are desirable catalysts for some purposes, but which may tend to produce polymer with LCB levels that are detrimental to film performance. Therefore, new catalyst compositions and methods that allow better control of LCB levels within a desired specification range is a desirable goal.

[0004] US 5,886,202 discloses bridged fluorenyl/indenyl metallocenes having substituents at the 3 position of the indenyl and the use thereof in the polymerization of olefins.

SUMMARY OF THE INVENTION

[0005] In the course of examining metallocene-based olefin polymerization catalysts, it was discovered that the long-chain branching (LCB) content of PE resins made with such catalysts was related to, among other things, the type of metallocene catalyst employed, and also related to the particular activator, including the particular solid oxide activator or "activator-support" which can constitute one component of the catalyst composition.

[0006] In one aspect of the present invention, for example, it was discovered that certain metallocene-based catalyst systems can produce high molecular weight polyethylene with low levels of LCB, even under relatively high temperature conditions. Useful metallocenes in preparing the catalyst compositions include tightly-bridged, *ansa*-metallocenes that comprise a pendant alkenyl (olefin-containing) group attached to at least one of the cyclopentadienyl-type moieties of the tightly-bridged ligand, and also comprises one or two aryl groups, particularly one or two phenyl groups, bonded to the bridging atom of the tightly-bridged ligand.

[0007] The subject matter of the invention is set out in the appended claims.

[0008] These and other features, aspects, embodiments, and advantages of the present invention will become apparent after a review of the following detailed description of the disclosed features.

BRIEF DESCRIPTION OF THE FIGURES

[0009]

FIGURE 1 illustrates the structures of specific metallocenes used in the inventive Examples.

FIGURE 2 illustrates the structures of specific metallocenes used in the comparative Examples.

FIGURE 3 illustrates data obtained from SEC-MALS analysis of the ethylene homopolymer produced in inventive Examples 1-4 (plot of R_g versus M_w).

FIGURE 4 illustrates data obtained from SEC-MALS analysis of the ethylene homopolymer produced in inventive Examples 5-7 (plot of R_g versus M_w).

FIGURE 5 illustrates data obtained from SEC-MALS analysis of the ethylene homopolymer produced in inventive Examples 10 and 11 (plot of R_g versus M_w).

FIGURE 6 provides a plot of zero shear viscosity versus molecular weight, specifically, $\log(\eta_0)$ versus $\log(M_w)$ for polymers prepared according to inventive Examples 1-11.

FIGURE 7 provides a plot of zero shear viscosity versus molecular weight, specifically, $\log(\eta_0)$ versus $\log(M_w)$ for polymers prepared according to comparative Examples 14-16.

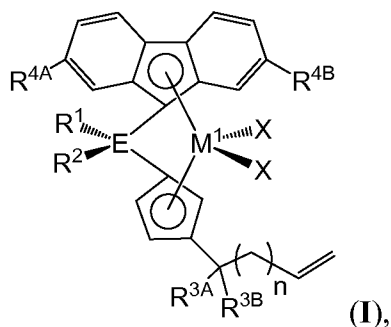
FIGURE 8 provides comparison gel permeation chromatograms (GPCs) for ethylene homopolymer of inventive Examples 1-11 (E1-E11) and comparative Examples 14-16 (E14-E16).

DETAILED DESCRIPTION OF THE INVENTION

Catalyst Composition and Components

The Metallocene Compounds

[0010] The at least one *ansa*-metallocene of this invention comprises a compound having the formula:



wherein:

M¹ is zirconium or hafnium;

X is, independently, F, Cl, Br, or I;

E is C;

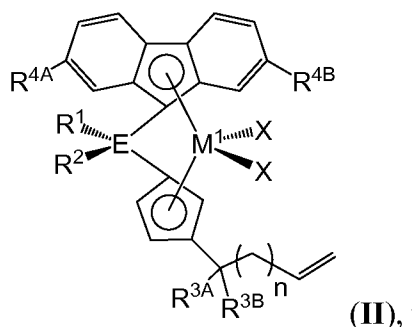
R¹ and R² are independently an aryl group, either of which having up to 10 carbon atoms;

R^{3A} and R^{3B} are independently a hydrocarbyl group or a trihydrocarbysilyl group, any of which having up to 20 carbon atoms; or hydrogen;

n is an integer from 0 to 10, inclusive; and

R^{4A} and R^{4B} are independently a hydrocarbyl group having up to 12 carbon atoms, or hydrogen;

[0011] In yet another aspect, the at least one *ansa*-metallocene of this invention comprises a compound having the formula:



wherein

M¹ is zirconium or hafnium;

X is F, Cl, Br, or I;

E is C;

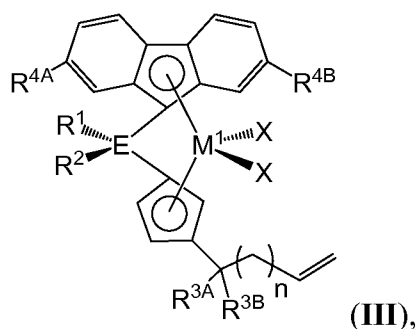
R¹ and R² are independently an aryl group, either of which having up to 10 carbon atoms;

R^{3A} and R^{3B} are independently H, methyl, allyl, benzyl, butyl, pentyl, hexyl, or trimethylsilyl;

n is an integer from 1 to 6, inclusive; and

R^{4A} and R^{4B} are independently a hydrocarbyl group having up to 6 carbon atoms, or hydrogen.

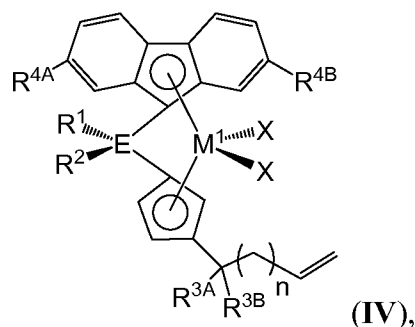
[0012] In still another aspect, the at least one *ansa*-metallocene of this invention comprises a compound having the formula:



wherein

- M¹ is zirconium or hafnium;
- X is Cl, Br, or I;
- E is C;
- R¹ and R² are independently phenyl;
- R^{3A} and R^{3B} are independently H or methyl;
- n is 1 or 2; and
- R^{4A} and R^{4B} are independently H or t-butyl.

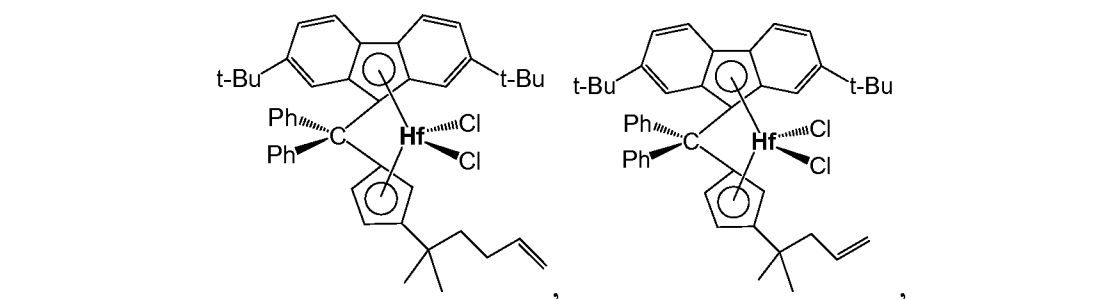
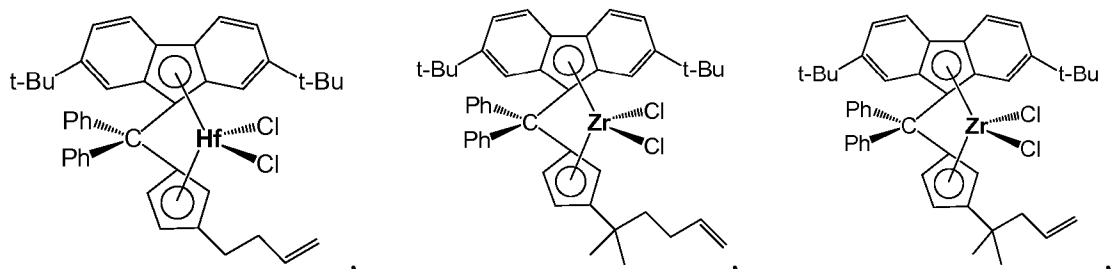
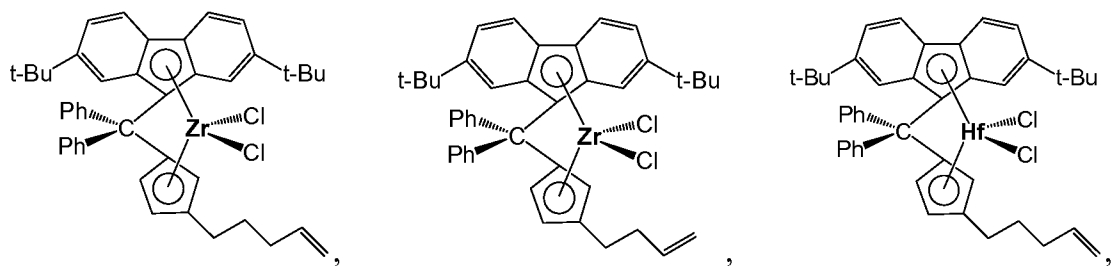
[0013] In yet another aspect, the at least one *ansa*-metallocene of this invention can comprise a compound having the formula:



wherein

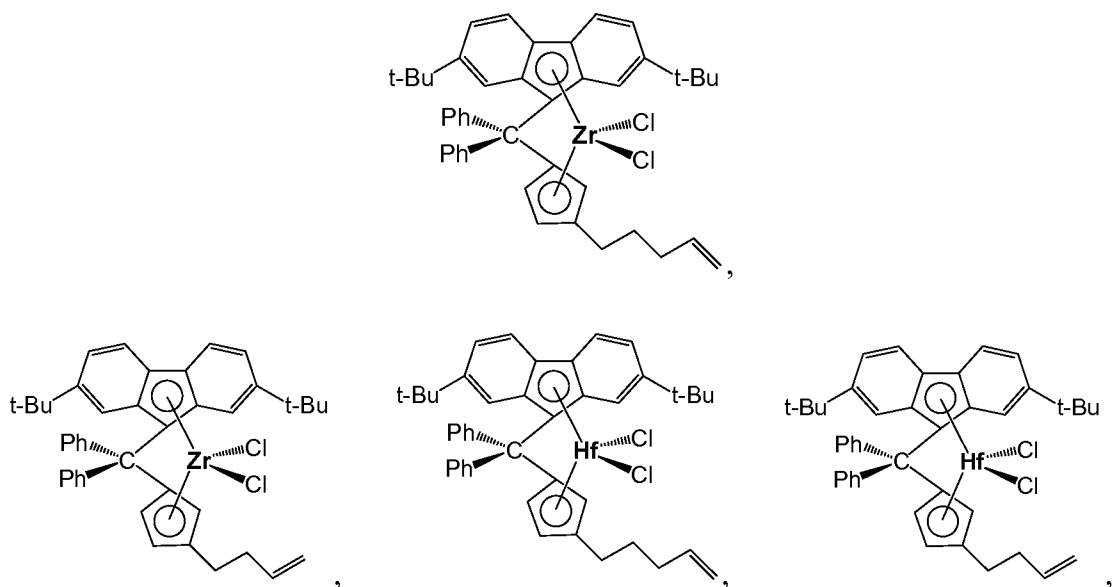
- M¹ is zirconium or hafnium;
- X is, independently, H, BH₄, methyl, phenyl, benzyl, neopentyl, trimethylsilylmethyl, CH₂CMe₂Ph; CH₂SiMe₂Ph; CH₂CMe₂CH₂Ph; or CH₂SiMe₂CH₂Ph;
- E is C;
- R¹ and R² are independently an aryl group, either of which having up to 10 carbon atoms;
- R^{3A} and R^{3B} are independently a hydrocarbyl group or a trihydrocarbylsilyl group, any of which having up to 20 carbon atoms; or hydrogen;
- n is an integer from 0 to 10, inclusive; and
- R^{4A} and R^{4B} are independently a hydrocarbyl group having up to 12 carbon atoms, or hydrogen.

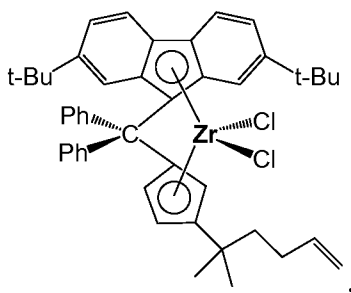
[0014] In a further aspect, the at least one *ansa*-metallocene of this invention comprises a compound having the formula:



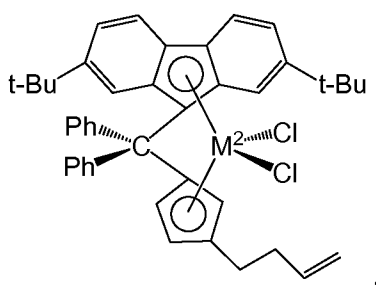
or any combination thereof.

[0015] In still another aspect, the at least one *ansa*-metallocene of this invention can comprise, or can be selected from,





or any combination thereof. Yet another aspect of the present invention provides a metallocene compound having the formula



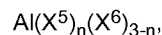
wherein M^2 is Zr or Hf.

[0016] Numerous processes to prepare metallocene compounds that can be employed in this invention have been reported. For example, U.S. Patent Nos. 4,939,217, 5,191,132, 5,210,352, 5,347,026, 5,399,636, 5,401,817, 5,420,320, 5,436,305, 5,451,649, 5,496,781, 5,498,581, 5,541,272, 5,554,795, 5,563,284, 5,565,592, 5,571,880, 5,594,078, 5,631,203, 5,631,335, 5,654,454, 5,668,230, 5,705,578, 5,705,579, 6,187,880, and 6,509,427 describe such methods. Other processes to prepare metallocene compounds that can be employed in this invention have been reported in references such as: Köppl, A. Alt, H. G. J. Mol. Catal. A. 2001, 165, 23; Kajigaeshi, S.; Kadowaki, T.; Nishida, A.; Fujisaki, S. The Chemical Society of Japan, 1986, 59, 97; Alt, H. G.; Jung, M.; Kehr, G. J. Organomet. Chem. 1998, 562, 153-181; Alt, H. G.; Jung, M. J. Organomet. Chem. 1998, 568, 87-112; Jung, M., Doctoral Dissertation, University of Bayreuth, Bayreuth, Germany, 1997; Piefer, B., Doctoral Dissertation, University of Bayreuth, Bayreuth, Germany, 1995; and Zenk, R., Doctoral Dissertation, University of Bayreuth, Bayreuth, Germany, 1994. The following treatises also describe such methods: Wailes, P. C.; Coutts, R. S. P.; Weigold, H. in Organometallic Chemistry of Titanium, Zirconium, and Hafnium, Academic; New York, 1974. ; Cardin, D. J.; Lappert, M. F.; and Raston, C. L.; Chemistry of Organo-Zirconium and -Hafnium Compounds; Halstead Press; New York, 1986.

The Organoaluminum Compound

[0017] There is also disclosed a catalyst composition comprising at least one compound according to the invention, at least one solid oxide activator-support, and optionally at least one organoaluminum compound. Thus, the designation of the at least one organoaluminum compound as optional is intended to reflect that the organoaluminum compound can be optional when it is not necessary to impart catalytic activity to the composition comprising the contact product, as understood by one of ordinary skill, as disclosed herein.

[0018] Organoaluminum compounds that can be used in this invention include compounds with the formula:



wherein (X^5) is a hydrocarbyl having from 1 to 20 carbon atoms; (X^6) is alkoxide or aryloxide, any of which having from 1 to 20 carbon atoms, halide, or hydride; and n is a number from 1 to 3, inclusive. In one aspect, (X^5) is an alkyl having from 1 to 10 carbon atoms. Examples of (X^5) moieties include methyl, ethyl, propyl, butyl, hexyl, heptyl, and octyl. In another aspect, examples of (X^5) moieties include methyl, ethyl, isopropyl, n-propyl, n-butyl, sec-butyl, isobutyl, 1-hexyl, 2-hexyl, 3-hexyl, isohexyl, heptyl, and octyl. In another aspect, (X^6) can be independently selected from fluoride, chloride, bromide, methoxide, ethoxide, or hydride. In yet another aspect, (X^6) can be chloride.

[0019] In the formula $Al(X^5)_n(X^6)_{3-n}$, n is a number from 1 to 3 inclusive, and typically, n is 3. The value of n is not

restricted to be an integer, therefore this formula includes sesquihalide compounds, and other organoaluminum cluster compounds.

[0020] Generally, examples of organoaluminum compounds that can be used in this invention include trialkylaluminum compounds, dialkylaluminum halide compounds, dialkylaluminum alkoxide compounds, dialkylaluminum hydride compounds, and combinations thereof. Examples of organoaluminum compounds that are useful in this invention include trimethylaluminum, triethylaluminum, tripropylaluminum, tributylaluminum, tri-n-butylaluminum (TNBA), triisobutylaluminum (TIBA), trihexylaluminum, triisohexylaluminum, trioctylaluminum, diethylaluminum ethoxide, diisobutylaluminum hydride, diethylaluminum chloride, or any combination thereof. If the particular alkyl isomer is not specified, the compound is intended to encompass all isomers that can arise from a particular specified alkyl group. Thus, in another aspect, examples of organoaluminum compounds that can be used in this invention include trimethylaluminum, triethylaluminum, tripropylaluminum, tributylaluminum, triisobutylaluminum, trihexylaluminum, triisohexylaluminum, trioctylaluminum, diethylaluminum ethoxide, diisobutylaluminum hydride, diethylaluminum chloride, or any combination thereof.

[0021] There is also disclosed precontacting the *ansa*-metallocene with at least one organoaluminum compound and an olefin monomer to form a precontacted mixture, prior to contact this precontacted mixture with the solid oxide activator-support to form the active catalyst. When the catalyst composition is prepared in this manner, typically, though not necessarily, a portion of the organoaluminum compound is added to the precontacted mixture and another portion of the organoaluminum compound is added to the postcontacted mixture prepared when the precontacted mixture is contacted with the solid oxide activator. However, all of the organoaluminum compound can be used to prepare the catalyst in either the precontacting or postcontacting step. Alternatively, all the catalyst components can be contacted in a single step.

[0022] Further, more than one organoaluminum compounds can be used, in either the precontacting or the postcontacting step, or in any procedure in which the catalyst components are contacted. When an organoaluminum compound is added in multiple steps, the amounts of organoaluminum compound disclosed herein include the total amount of organoaluminum compound used in both the precontacted and postcontacted mixtures, and any additional organoaluminum compound added to the polymerization reactor. Therefore, total amounts of organoaluminum compounds are disclosed, regardless of whether a single organoaluminum compound is used, or more than one organoaluminum compound. In another aspect, typical organoaluminum compounds used in this invention include triethylaluminum (TEA), tri-n-butylaluminum, triisobutylaluminum, or any combination thereof.

The Activator

[0023] There is also disclosed a catalyst composition comprising at least one tightly-bridged *ansa*-metallocene compound as disclosed herein; optionally, at least one organoaluminum compound; and at least one activator. In another aspect, the at least one activator can be an activator-support comprising a solid oxide treated with an electron-withdrawing anion; a layered mineral; an ion-exchangeable activator-support; an organoaluminum compound; an organoboron compound; an organoborate compound; or any combination of any of these activators, each of which is provided herein.

Chemically-Treated Solid Oxide Activator-Supports

[0024] There is also disclosed catalyst compositions comprising an acidic activator-support, which can comprise a chemically-treated solid oxide, and which is typically used in combination with an organoaluminum compound. In another aspect, the activator-support comprises at least one solid oxide treated with at least one electron-withdrawing anion; wherein the solid oxide can be silica, alumina, silica-alumina, aluminophosphate, aluminum phosphate, zinc aluminate, heteropolytungstates, titania, zirconia, magnesia, boria, zinc oxide, mixed oxides thereof or any mixture or combination thereof; and wherein the electron-withdrawing anion can be fluoride, chloride, bromide, iodide, phosphate, triflate, bisulfate, sulfate, fluoroborate, fluorosulfate, trifluoroacetate, phosphate, fluorophosphate, fluorozirconate, fluorosilicate, fluorotitanate, permanganate, substituted or unsubstituted alkanesulfonate, substituted or unsubstituted arenesulfonate, substituted or unsubstituted alkylsulfate, or any combination thereof.

[0025] The activator-support includes the contact product of at least one solid oxide compound and at least one electron-withdrawing anion source. In one aspect, the solid oxide compound comprises an inorganic oxide. The solid oxide can be optionally calcined prior to contacting the electron-withdrawing anion source. The contact product can also be calcined either during or after the solid oxide compound is contacted with the electron-withdrawing anion source. In this aspect, the solid oxide compound can be calcined or uncalcined. In another aspect, the activator-support can comprise the contact product of at least one calcined solid oxide compound and at least one electron-withdrawing anion source.

[0026] The activator-support exhibits enhanced activity as compared to the corresponding untreated solid oxide compound. The activator-support also functions as a catalyst activator as compared to the corresponding untreated solid oxide. While not intending to be bound by theory, it is believed that the activator-support can function as solid oxide

supporting compound with an additional ionizing, polarizing, or bond weakening function, collectively termed an "activating" function, by weakening the metal-ligand bond between an anionic ligand and the metal in the metallocene. Thus, the activator-support is considered to exhibit an "activating" function, regardless of whether it ionizes the metallocene, abstracts an anionic ligand to form an ion pair, weakens the metal-ligand bond in the metallocene, simply coordinates to an anionic ligand when it contacts the activator-support, or any other mechanisms by which ionizing, polarizing, or bond weakening might occur. In preparing the metallocene-based catalyst composition, the activator-support is typically used along with a component that provides an activatable ligand such as an alkyl or hydride ligand to the metallocene, including at least one organoaluminum compound, when the metallocene compound does not already comprise such a ligand.

[0027] The activator-support may comprise a solid inorganic oxide material, a mixed oxide material, or a combination of inorganic oxide materials, that is chemically-treated with an electron-withdrawing component, and optionally treated with at least one other metal ion. Thus, the solid oxide encompasses oxide materials such as alumina, "mixed oxide" compounds such as silica-alumina or silica-zirconia or silica-titania, and combinations and mixtures thereof. The mixed metal oxide compounds such as silica-alumina, with more than one metal combined with oxygen to form a solid oxide compound, can be made by co-gellation, impregnation or chemical deposition.

[0028] The activator-support may further comprise a metal or metal ion such as zinc, nickel, vanadium, silver, copper, gallium, tin, tungsten, molybdenum, or any combination thereof. Examples of activator-supports that further comprise a metal or metal ion include zinc-impregnated chlorided alumina, zinc-impregnated fluorided alumina, zinc-impregnated chlorided silica-alumina, zinc-impregnated fluorided silica-alumina, zinc-impregnated sulfated alumina, or any combination thereof.

[0029] The activator-support may comprise a solid oxide of relatively high porosity, which exhibits Lewis acidic or Brønsted acidic behavior. The solid oxide is chemically-treated with an electron-withdrawing component, typically an electron-withdrawing anion, to form a activator-support. While not intending to be bound by the following statement, it is believed that treatment of the inorganic oxide with an electron-withdrawing component augments or enhances the acidity of the oxide. Thus, the activator-support exhibits Lewis or Brønsted acidity which is typically greater than the Lewis or Brønsted acidity of the untreated solid oxide. One method to quantify the acidity of the chemically-treated and untreated solid oxide materials is by comparing the polymerization activities of the treated and untreated oxides under acid catalyzed reactions.

[0030] The chemically-treated solid oxide may comprise a solid inorganic oxide comprising oxygen and at least one element selected from Group 2, 3, 4, 5, 6, 7, 8, 9, 10, 11, 12, 13, 14, or 15 of the periodic table, or comprising oxygen and at least one element selected from the lanthanide or actinide elements. (See: Hawley's Condensed Chemical Dictionary, 11th Ed., John Wiley & Sons; 1995; Cotton, F.A.; Wilkinson, G.; Murillo, C. A.; and Bochmann, M. Advanced Inorganic Chemistry, 6th Ed., Wiley-Interscience, 1999.) Usually, the inorganic oxide comprises oxygen and at least one element selected from Al, B, Be, Bi, Cd, Co, Cr, Cu, Fe, Ga, La, Mn, Mo, Ni, Sb, Si, Sn, Sr, Th, Ti, V, W, P, Y, Zn or Zr.

[0031] Suitable examples of solid oxide materials or compounds that can be used in the chemically-treated solid oxide include Al_2O_3 , B_2O_3 , BeO, Bi_2O_3 , CdO, Co_3O_4 , Cr_2O_3 , CuO, Fe_2O_3 , Ga_2O_3 , La_2O_3 , Mn_2O_3 , MoO_3 , NiO, P_2O_5 , Sb_2O_5 , SiO_2 , SnO_2 , SrO, ThO_2 , TiO_2 , V_2O_5 , WO_3 , Y_2O_3 , ZnO, ZrO_2 , including mixed oxides thereof, and combinations thereof. Examples of mixed oxides that can be used in the activator-support of the present invention include mixed oxides of any combination of Al, B, Be, Bi, Cd, Co, Cr, Cu, Fe, Ga, La, Mn, Mo, Ni, P, Sb, Si, Sn, Sr, Th, Ti, V, W, Y, Zn, Zr. Examples of mixed oxides that can be used in the activator-support of the present invention also include silica-alumina, silica-titania, silica-zirconia, zeolites, many clay minerals, pillared clays, alumina-titania, alumina-zirconia, and aluminophosphate.

[0032] The solid oxide material may be chemically-treated by contacting it with at least one electron-withdrawing component, typically an electron-withdrawing anion source. Further, the solid oxide material is optionally chemically-treated with at least one other metal ion, that can be the same as or different from any metal element that constitutes the solid oxide material, then calcining to form a metal-containing or metal-impregnated chemically-treated solid oxide. Alternatively, a solid oxide material and an electron-withdrawing anion source are contacted and calcined simultaneously. The method by which the oxide is contacted with an electron-withdrawing component, typically a salt or an acid of an electron-withdrawing anion, includes gelling, co-gelling, impregnation of one compound onto another. Typically, following any contacting method, the contacted mixture of oxide compound, electron-withdrawing anion, and optionally the metal ion is calcined.

[0033] The electron-withdrawing component used to treat the oxide can be any component that increases the Lewis or Brønsted acidity of the solid oxide upon treatment. The electron-withdrawing component is typically an electron-withdrawing anion derived from a salt, an acid, or other compound such as a volatile organic compound that can serve as a source or precursor for that anion. Examples of electron-withdrawing anions include fluoride, chloride, bromide, iodide, phosphate, triflate, bisulfate, sulfate, fluoroborate, fluorosulfate, trifluoroacetate, phosphate, fluorophosphate, fluorozirconate, fluorosilicate, fluorotitanate, permanganate, substituted or unsubstituted alkanesulfonate, substituted or unsubstituted arenesulfonate, substituted or unsubstituted alkylsulfate, including any mixtures and combinations

thereof. In addition, other ionic or non-ionic compounds that serve as sources for these electron-withdrawing anions can also be employed in the present invention. The chemically-treated solid oxide may comprise a sulfated solid oxide, and in another aspect, the chemically-treated oxide comprises sulfated alumina.

[0034] The terms alkanesulfonate and alkyl sulfate refer to anions having the general formula $[R^B\text{SO}_2\text{O}]^-$ and $[(R^B\text{O})\text{SO}_2\text{O}]^-$, respectively, wherein R^B is a linear or branched alkyl group having up to 20 carbon atoms, that is optionally substituted with at least one group selected independently from F, Cl, Br, I, OH, OMe, OEt, OCF_3 , Ph, xylyl, mesityl, or OPh. Thus, the alkanesulfonate and alkyl sulfate may be referred to as being either substituted or unsubstituted. In one aspect, the alkyl group of the alkanesulfonate or alkyl sulfate can have up to 12 carbon atoms. In another aspect, the alkyl group of the alkanesulfonate or alkyl sulfate can have up to 8 carbon atoms, or up to 6 carbon atoms. In yet another aspect, examples of alkanesulfonates include, methanesulfonate, ethanesulfonate, 1-propanesulfonate, 2-propanesulfonate, 3-methylbutanesulfonate, trifluoromethanesulfonate, trichloromethanesulfonate, chloromethanesulfonate, 1-hydroxyethanesulfonate, 2-hydroxy-2-propanesulfonate, and 1-methoxy-2-propanesulfonate. In still another aspect, examples of alkyl sulfates include methylsulfate, ethylsulfate, 1-propylsulfate, 2-propylsulfate, 3-methylbutylsulfate, trifluoromethanesulfate, trichloromethylsulfate, chloromethylsulfate, 1-hydroxyethylsulfate, 2-hydroxy-2-propylsulfate, and 1-methoxy-2-propylsulfate.

[0035] The term arenesulfonate refers to anions having the general formula $[\text{Ar}^A\text{SO}_2\text{O}]^-$, wherein Ar^A is an aryl group having up to 14 carbon atoms, that is optionally substituted with at least one group selected independently from F, Cl, Br, I, Me, Et, Pr, Bu, OH, OMe, OEt, OPr, OBu, OCF_3 , Ph, OPh, or R^C , wherein R^C is a linear or branched alkyl group having up to 20 carbon atoms. Thus, the arenesulfonate may be referred to as a substituted or an unsubstituted arenesulfonate. Because the aryl group Ar^A can be substituted with an alkyl side chain, R^C , which includes a long alkyl side chain, the term arenesulfonate is intended to include the detergents. The aryl group of the arenesulfonate can have up to 10 carbon atoms. In another aspect, the aryl group of the arenesulfonate can have 6 carbon atoms. In yet another aspect, examples of arenesulfonates include benzenesulfonate, naphthalenesulfonate, *p*-toluenesulfonate, *m*-toluenesulfonate, 3,5-xylenesulfonate, trifluoromethoxybenzenesulfonate, trichloromethoxybenzenesulfonate, trifluoromethylbenzenesulfonate, trichloromethylbenzenesulfonate, fluorobenzenesulfonate, chlorobenzenesulfonate, 1-hydroxyethanebenzenesulfonate, and 3-fluoro-4-methoxybenzenesulfonate.

[0036] When the electron-withdrawing component comprises a salt of an electron-withdrawing anion, the counterion or cation of that salt can be any cation that allows the salt to revert or decompose back to the acid during calcining. Factors that dictate the suitability of the particular salt to serve as a source for the electron-withdrawing anion include the solubility of the salt in the desired solvent, the lack of adverse reactivity of the cation, ion-pairing effects between the cation and anion, hygroscopic properties imparted to the salt by the cation and thermal stability of the anion. Examples of suitable cations in the salt of the electron-withdrawing anion include ammonium, trialkyl ammonium, tetraalkyl ammonium, tetraalkyl phosphonium, H^+ , and $[\text{H}(\text{OEt}_2)_2]^+$.

[0037] Further, combinations of one or more different electron withdrawing anions, in varying proportions, can be used to tailor the specific acidity of the activator-support to the desired level. Combinations of electron withdrawing components can be contacted with the oxide material simultaneously or individually, and any order that affords the desired activator-support acidity. For example, two or more electron-withdrawing anion source compounds in two or more separate contacting steps can be employed. Thus, one example of such a process by which an activator-support is prepared is as follows. A selected solid oxide compound, or combination of oxide compounds, is contacted with a first electron-withdrawing anion source compound to form a first mixture, this first mixture is then calcined, the calcined first mixture is then contacted with a second electron-withdrawing anion source compound to form a second mixture, followed by calcining said second mixture to form a treated solid oxide compound. In such a process, the first and second electron-withdrawing anion source compounds are typically different compounds, although they can be the same compound.

[0038] The solid oxide activator-support may be produced by a process comprising:

- 1) contacting a solid oxide compound with at least one electron-withdrawing anion source compound to form a first mixture; and
- 2) calcining the first mixture to form the solid oxide activator-support.

[0039] The solid oxide activator-support may be produced by a process comprising:

- 1) contacting at least one solid oxide compound with a first electron-withdrawing anion source compound to form a first mixture; and
- 2) calcining the first mixture to produce a calcined first mixture;
- 3) contacting the calcined first mixture with a second electron-withdrawing anion source compound to form a second mixture; and
- 4) calcining the second mixture to form the solid oxide activator-support.

Thus, the solid oxide activator-support is sometimes referred to simply as a treated solid oxide compound.

[0040] Producing or forming the solid oxide activator-support may be by contacting at least one solid oxide with at least one electron-withdrawing anion source compound, wherein the at least one solid oxide compound is calcined before, during or after contacting the electron-withdrawing anion source, and wherein there is a substantial absence of aluminoxanes and organoborates.

[0041] Once the solid oxide has been treated and dried, it can be subsequently calcined. Calcining of the treated solid oxide is generally conducted in an ambient or inert atmosphere, typically in a dry ambient atmosphere, at a temperature from 200°C to 900°C, and for a time of 1 minute to 100 hours. In another aspect, calcining is conducted at a temperature from 300°C to 800°C and in another aspect, calcining is conducted at a temperature from 400°C to 700°C. In yet another aspect, calcining is conducted from 1 hour to 50 hours, and in another aspect calcining is conducted, from 3 hours to 20 hours. In still another aspect, calcining can be carried out from 1 to 10 hours at a temperature from 350°C to 550°C.

[0042] Further, any type of suitable ambient can be used during calcining. Generally, calcining is conducted in an oxidizing atmosphere, such as air. Alternatively, an inert atmosphere, such as nitrogen or argon, or a reducing atmosphere such as hydrogen or carbon monoxide, can be used.

[0043] The solid oxide component used to prepare the chemically-treated solid oxide may have a pore volume greater than 0.1 cc/g. The solid oxide component may have a pore volume greater than 0.5 cc/g, and greater than 1.0 cc/g. The solid oxide component may have a surface area from 100 to 1000 m²/g. The solid oxide component may have a surface area from 200 to 800 m²/g, and in still another aspect, from 250 to 600 m²/g.

[0044] The solid oxide material can be treated with a source of halide ion or sulfate ion, or a combination of anions, and optionally treated with at least one metal ion, then calcined to provide the activator-support in the form of a particulate solid. The solid oxide material may be treated with a source of sulfate, termed a sulfating agent, a source of chloride ion, termed a chloriding agent, a source of fluoride ion, termed a fluoriding agent, or a combination thereof, and calcined to provide the solid oxide activator. Useful acidic activator-supports include: bromided alumina; chlorided alumina; fluorided alumina; sulfated alumina; bromided silica-alumina, chlorided silica-alumina; fluorided silica-alumina; sulfated silica-alumina; bromided silica-zirconia, chlorided silica-zirconia; fluorided silica-zirconia; sulfated silica-zirconia; chlorided zinc-alumina, triflate treated silica-alumina, a pillared clay such as a pillared montmorillonite, optionally treated with fluoride, chloride, or sulfate; phosphated alumina, or other aluminophosphates, optionally treated with sulfate, fluoride, or chloride; or any combination thereof. Further, any of the activator-supports can optionally be treated with at least one other metal ion, typically from a metal salt or compound, wherein the metal ion can be the same as or different from any metal that makes up the solid oxide material.

[0045] The treated oxide activator-support may comprise a fluorided solid oxide in the form of a particulate solid, thus a source of fluoride ion is added to the oxide by treatment with a fluoriding agent. Fluoride ion can be added to the oxide by forming a slurry of the oxide in a suitable solvent such as alcohol or water, including the one to three carbon alcohols because of their volatility and low surface tension. Examples of fluoriding agents that can be used in this invention include hydrofluoric acid (HF), ammonium fluoride (NH₄F), ammonium bifluoride (NH₄HF₂), ammonium tetrafluoroborate (NH₄BF₄), ammonium silicofluoride (hexafluorosilicate) ((NH₄)₂SiF₆), ammonium hexafluorophosphate (NH₄PF₆), tetrafluoroboric acid (HBF₄), ammonium hexafluorotitanate (NH₄)₂TiF₆, ammonium hexafluorozirconate (NH₄)₂ZrF₆, analogs thereof, and combinations thereof. For example, ammonium bifluoride NH₄HF₂ can be used as the fluoriding agent, due to its ease of use and ready availability.

[0046] The solid oxide can be treated with a fluoriding agent during the calcining step. Any fluoriding agent capable of thoroughly contacting the solid oxide during the calcining step can be used. For example, in addition to those fluoriding agents described previously, volatile organic fluoriding agents can be used. Examples of volatile organic fluoriding agents useful in this aspect of the invention include freons, perfluorohexane, perfluorobenzene, fluoromethane, trifluoroethanol, and combinations thereof. Gaseous hydrogen fluoride or fluorine itself can also be used with the solid oxide is fluorided during calcining. One convenient method of contacting the solid oxide with the fluoriding agent is to vaporize a fluoriding agent into a gas stream used to fluidize the solid oxide during calcination.

[0047] Similarly, the chemically-treated solid oxide may comprise a chlorided solid oxide in the form of a particulate solid, thus a source of chloride ion is added to the oxide by treatment with a chloriding agent. The chloride ion can be added to the oxide by forming a slurry of the oxide in a suitable solvent. The solid oxide can be treated with a chloriding agent during the calcining step. Any chloriding agent capable of serving as a source of chloride and thoroughly contacting the oxide during the calcining step can be used. For example, volatile organic chloriding agents can be used. Examples of volatile organic chloriding agents useful in this aspect of the invention include certain freons, perchlorobenzene, chloromethane, dichloromethane, chloroform, carbon tetrachloride, trichloroethanol, or any combination thereof. Gaseous hydrogen chloride or chlorine itself can also be used with the solid oxide during calcining. One convenient method of contacting the oxide with the chloriding agent is to vaporize a chloriding agent into a gas stream used to fluidize the solid oxide during calcination.

[0048] When the activator-support comprises a chemically-treated solid oxide comprising a solid oxide treated with an electron-withdrawing anion, the electron withdrawing anion can be typically added to the solid oxide in an amount

greater than 1% by weight of the solid oxide. The electron withdrawing anion can be added to the solid oxide in an amount greater than 2% by weight of the solid oxide, greater than 3% by weight of the solid oxide, greater than 5% by weight of the solid oxide, or greater than 7% by weight of the solid oxide.

[0049] The amount of electron-withdrawing ion, for example fluoride or chloride ion, present before calcining the solid oxide is generally from 2 to 50% by weight, where the weight percents are based on the weight of the solid oxide, for example silica-alumina, before calcining. The amount of electron-withdrawing ion, for example fluoride or chloride ion, present before calcining the solid oxide is from 3 to 25% by weight, and from 4 to 20% by weight. When halide ion is used as the electron-withdrawing anion, it is used in an amount sufficient to deposit, after calcining, from 0.1% to 50% by weight halide ion relative to the weight of the solid oxide. In another aspect, halide is used in an amount sufficient to deposit, after calcining, from 0.5% to 40% by weight halide ion relative to the weight of the solid oxide, or from 1% to 30% by weight halide ion relative to the weight of the solid oxide. If the fluoride or chloride ion is added during calcining, such as when calcined in the presence of CCl_4 , there is typically no, or only trace levels, of fluoride or chloride ion in the solid oxide before calcining. Once impregnated with halide, the halided oxide can be dried by any method known in the art including suction filtration followed by evaporation, drying under vacuum, and spray drying, although it is also possible to initiate the calcining step immediately without drying the impregnated solid oxide.

[0050] The silica-alumina used to prepare the treated silica-alumina can have a pore volume greater than 0.5 cc/g. The pore volume can be greater than 0.8 cc/g, and the pore volume can be greater than 1.0 cc/g. Further, the silica-alumina can have a surface area greater than 100 m^2/g . The surface area may be greater than 250 m^2/g , and the surface area can be greater than 350 m^2/g . Generally, the silica-alumina of this invention has an alumina content from 5 to 95%. The alumina content of the silica-alumina can be from 5 to 50%, and the alumina content of the silica-alumina can be from 8% to 30% alumina by weight.

[0051] The sulfated solid oxide comprises sulfate and a solid oxide component such as alumina or silica-alumina, in the form of a particulate solid. Optionally, the sulfated oxide is further treated with a metal ion such that the calcined sulfated oxide comprises a metal. The sulfated solid oxide may comprise sulfate and alumina. The sulfated alumina may be formed by a process wherein the alumina is treated with a sulfate source, including for example sulfuric acid or a sulfate salt such as ammonium sulfate, zinc sulfate, aluminum sulfate, nickel sulfate or copper sulfate. This process can be performed by forming a slurry of the alumina in a suitable solvent such as alcohol or water, in which the desired concentration of the sulfating agent has been added. Suitable organic solvents include the one to three carbon alcohols because of their volatility and low surface tension.

[0052] The amount of sulfate ion present before calcining is generally from 1% to 50% by weight, from 2% to 30% by weight, of from 5% to 25% by weight, where the weight percents are based on the weight of the solid oxide before calcining. Once impregnated with sulfate, the sulfated oxide can be dried by any method known in the art including suction filtration followed by evaporation, drying under vacuum, and spray drying, although it is also possible to initiate the calcining step immediately.

[0053] In addition to being treated with an electron-withdrawing component such as halide or sulfate ion, the solid inorganic oxide can optionally be treated with a metal source, including metal salts or metal-containing compounds. These compounds can be added to or impregnated onto the solid oxide in solution form, and subsequently converted into the supported metal upon calcining. Accordingly, the solid inorganic oxide can further comprise a metal selected from zinc, nickel, vanadium, silver, copper, gallium, tin, tungsten, molybdenum, or a combination thereof. For example, zinc can be used to impregnate the solid oxide because it provides good catalyst activity and low cost. The solid oxide can be treated with metal salts or metal-containing compounds before, after, or at the same time that the solid oxide is treated with the electron-withdrawing anion.

[0054] Further, any method of impregnating the solid oxide material with a metal can be used. The method by which the oxide is contacted with a metal source, typically a salt or metal-containing compound, includes gelling, co-gelling, and impregnation of one compound onto another. Following any contacting method, the contacted mixture of oxide compound, electron-withdrawing anion, and the metal ion is typically calcined. Alternatively, a solid oxide material, an electron-withdrawing anion source, and the metal salt or metal-containing compound are contacted and calcined simultaneously.

[0055] The *ansa*-metallocene compound can be contacted with an olefin monomer and an organoaluminum cocatalyst for a first period of time prior to contacting this mixture with the acidic activator-support. Once the precontacted mixture of metallocene, monomer, and component that provides an activatable ligand to the metallocene, including an organoaluminum cocatalyst, is contacted with the acidic activator-support, the composition further comprising the acidic activator-support is termed the "postcontacted" mixture. The postcontacted mixture can be allowed to remain in further contact for a second period of time prior to being charged into the reactor in which the polymerization process will be carried out.

[0056] Various processes to prepare solid oxide activator-supports that can be employed in this invention have been reported. For example, U.S. Patent Nos. 6,107,230, 6,165,929, 6,294,494, 6,300,271, 6,316,553, 6,355,594, 6,376,415, 6,391,816, 6,395,666, 6,524,987, and 6,548,441, describe such methods.

Ion-Exchangeable Activator-Supports and Layered Mineral Activator-Supports

[0057] The activator-support used in preparing the catalyst compositions can comprise an ion-exchangeable activator-support, including silicate and aluminosilicate compounds or minerals, either with layered or non-layered structures, and any combination thereof. In another aspect of this invention, ion-exchangeable, layered aluminosilicates such as pillared clays can be used as activator-supports. When the acidic activator-support comprises an ion-exchangeable activator-support, it can optionally be treated with at least one electron-withdrawing anion such as those disclosed herein, though typically the ion-exchangeable activator-support is not treated with an electron-withdrawing anion.

[0058] The activator-support can comprise clay minerals having exchangeable cations and layers capable of expanding. Typical clay mineral activator-supports include ion-exchangeable, layered aluminosilicates such as pillared clays. Although the term "support" is used, it is not meant to be construed as an inert component of the catalyst composition, but rather is to be considered an active part of the catalyst composition, because of its intimate association with the *ansa*-metallocene and the component that provides an activatable ligand to the metallocene, such as an organoaluminum. While not intending to be bound by theory, it is believed that the ion exchangeable activator-support serves as an insoluble reactant that reacts with the *ansa*-metallocene and organoaluminum components to form a catalyst composition used to produce polymer.

[0059] The clay materials encompass materials either in their natural state or that have been treated with various ions by wetting, ion exchange, or pillaring. Typically, the clay material activator-support of this invention comprises clays that have been ion exchanged with large cations, including polynuclear, highly charged metal complex cations. However, the clay material activator-supports of this invention also encompass clays that have been ion exchanged with simple salts, including, but not limited to, salts of Al(III), Fe(II), Fe(III) and Zn(II) with ligands such as halide, acetate, sulfate, nitrate, or nitrite.

[0060] The clay activator-support comprises pillared clays. The term pillared clays is used to refer to clay materials that have been ion exchanged with large, typically polynuclear, highly charged metal complex cations. Examples of such ions include Keggin ions which can have charges such as 7+, various polyoxometallates, and other large ions. Thus, the term pillaring refers to a simple exchange reaction in which the exchangeable cations of a clay material are replaced with large, highly charged ions, such as Keggin ions. These polymeric cations are then immobilized within the interlayers of the clay and when calcined are converted to metal oxide "pillars," effectively supporting the clay layers as column-like structures. Thus, once the clay is dried and calcined to produce the supporting pillars between clay layers, the expanded lattice structure is maintained and the porosity is enhanced. The resulting pores can vary in shape and size as a function of the pillaring material and the parent clay material used. Examples of pillaring and pillared clays are found in: T.J. Pinnavaia, *Science* 220 (4595), 365-371 (1983); J.M. Thomas, *Intercalation Chemistry*, (S. Whittington and A. Jacobson, eds.) Ch. 3, pp. 55-99, Academic Press, Inc., (1972); U.S. Patent No. 4,452,910; U.S. Patent No. 5,376,611; and U.S. Patent No. 4,060,480.

[0061] The pillaring process utilizes clay minerals having exchangeable cations and layers capable of expanding. Any pillared clay that can enhance the polymerization of olefins in the catalyst composition of the present invention can be used. Therefore, suitable clay minerals for pillaring include: allophanes; smectites, both dioctahedral (Al) and tri-octahedral (Mg) and derivatives thereof such as montmorillonites (bentonites), nontronites, hectorites, or laponites; halloysites; vermiculites; micas; fluoromicas; chlorites; mixed-layer clays; the fibrous clays including sepiolites, attapulgites, and palygorskites; a serpentine clay; illite; laponite; saponite; or any combination thereof. The pillared clay activator-support may comprise bentonite or montmorillonite, noting that the principal component of bentonite is montmorillonite.

[0062] The pillared clay can be pretreated. For example, in one embodiment, a pillared bentonite was pretreated by drying at about 300°C under an inert atmosphere, typically dry nitrogen, for about 3 hours, before being added to the polymerization reactor. This example of a pretreatment is not limiting, because preheating steps such as this may be carried out at many other temperatures and times, including a combination of temperature and time steps, all of which are encompassed.

[0063] The ion-exchangeable activator-supports such as pillared clays used to prepare the catalyst compositions can be combined with other inorganic support materials, including zeolites, inorganic oxides, and phosphated inorganic oxides. Typical support materials that can be used in this regard include silica, silica-alumina, alumina, titania, zirconia, magnesia, boria, fluorided alumina, sited alumina, thoria, aluminophosphate, aluminum phosphate, zinc aluminate, phosphated silica, phosphated alumina, silica-titania, coprecipitated silica/titania, fluorided/sited alumina, and any combination or mixture thereof.

[0064] The amount of *ansa*-metallocene compound in relation to the ion-exchangeable activator-support used to prepare the catalyst composition is typically from 0.1 wt% to 15 wt% *ansa*-metallocene complex, based on the weight of the activator-support component (not based on the final metallocene-clay mixture). It was also found that from 1 wt% to 10 wt% *ansa*-metallocene works well to afford a catalyst that operates at desired activities.

[0065] The mixture of *ansa*-metallocene and clay activator-support can be contacted and mixed for any length of time to allow thorough contact between the *ansa*-metallocene and activator-support. Sufficient deposition of the metallocene

component on the clay can be achieved without heating a mixture of clay and metallocene complex. For example, the *ansa*-metallocene compound and the clay material are simply mixed from about room temperature to 93°C (200°F) in order to achieve the deposition of the *ansa*-metallocene on the clay activator-support. The *ansa*-metallocene compound and the clay material may be mixed from 37.8°C to 82.2°C (100°F to 180°F) in order to achieve the deposition of the *ansa*-metallocene on the clay activator-support.

[0066] There is also disclosed catalyst compositions comprising an acidic activator-support, which can comprise a layered mineral. The term "layered mineral" is used herein to describe materials such as clay minerals, pillared clays, ion-exchanged clays, exfoliated clays, exfoliated clays gelled into another oxide matrix, layered minerals mixed or diluted with other materials, or any combination thereof. When the acidic activator-support comprises a layered mineral, it can optionally be treated with at least one electron-withdrawing anion such as those disclosed herein, though typically the layered mineral is not treated with an electron-withdrawing anion. For example, at least one clay mineral can be used as the activator-support.

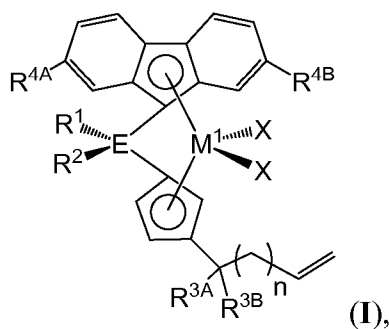
[0067] Clay minerals generally include the large group of finely-crystalline, sheet-like layered minerals that are found in nature in fine-grained sediments, sedimentary rocks and which constitute a class of hydrous silicate and aluminosilicate minerals with sheet-like structures and very high surface areas. This term is also used to describe hydrous magnesium silicates with a phyllosilicate structure. Examples of clay minerals that can be used in this invention include allophanes; smectites, both dioctahedral (Al) and tri-octahedral (Mg) and derivatives thereof such as montmorillonites (bentonites), nontronites, hectorites, or laponites; halloysites; vermiculites; micas; fluoromicas; chlorites; mixed-layer clays; the fibrous clays including sepiolites, attapulgites, and palygorskites; a serpentine clay; illite; laponite; saponite; or any combination thereof. Many common clay minerals belong to the kaolinite, montmorillonite, or illite groups of clays. Pillared clays can also be used as the activator-support of this invention, as disclosed herein. Pillared clays comprise clay minerals, typically of the of the smectite group and other phyllosilicates in addition to sepiolites and palygorskites, that have been ion exchanged with large, typically polynuclear, highly charged metal complex cations.

[0068] When layered minerals are used as activator-supports or metallocene activators, the layered minerals are typically calcined prior to their use as activators. Typical calcination temperatures can range from 100°C to 700°C, from 150°C to 500°C, or from 200°C to 400°C.

Non-Limiting Examples of the Catalyst Composition

[0069] Examples of the catalyst composition include the following. The catalyst composition can comprise, or the catalyst composition can comprise the contact product of, at least one *ansa*-metallocene, at least one organoaluminum compound, and at least one activator-support, wherein:

a) the at least one *ansa*-metallocene comprises a compound having the formula:



wherein

M¹ is zirconium or hafnium;

X is, independently, F, Cl, Br, or I;

E is C;

R¹ and R² are independently an aryl group, either of which having up to 10 carbon atoms;

R^{3A} and R^{3B} are independently a hydrocarbyl group or a trihydrocarbysilyl group, any of which having up to 20 carbon atoms; or hydrogen;

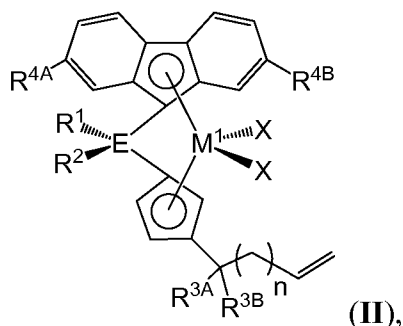
n is an integer from 0 to 10; and

R^{4A} and R^{4B} are independently a hydrocarbyl group having up to 12 carbon atoms, or hydrogen

- b) the at least one organoaluminum compound comprises trimethylaluminum, triethylaluminum, tripropylaluminum, tributylaluminum, triisobutylaluminum, trihexylaluminum, triisohexylaluminum, trioctylaluminum, diethylaluminum ethoxide, diisobutylaluminum hydride, diethylaluminum chloride, or any combination thereof; and
- c) the at least one activator-support comprises a solid oxide treated with an electron-withdrawing anion, wherein

the solid oxide is silica, alumina, silica-alumina, aluminophosphate, aluminum phosphate, zinc aluminate, heteropolytungstates, titania, zirconia, magnesia, boria, zinc oxide, mixed oxides thereof, or any combination thereof; and the electron-withdrawing anion is fluoride, chloride, bromide, iodide, phosphate, triflate, bisulfate, sulfate, fluoroborate, fluorosulfate, trifluoroacetate, phosphate, fluorophosphate, fluorozirconate, fluorosilicate, fluorotitanate, permanganate, substituted or unsubstituted alkanesulfonate, substituted or unsubstituted arenesulfonate, substituted or unsubstituted alkylsulfate, or any combination thereof.

[0070] The at least one *ansa*-metallocene can also comprise, or can be selected from, a compound having the formula:



wherein

M¹ is zirconium or hafnium;

X is F, Cl, Br, or I;

E is C;

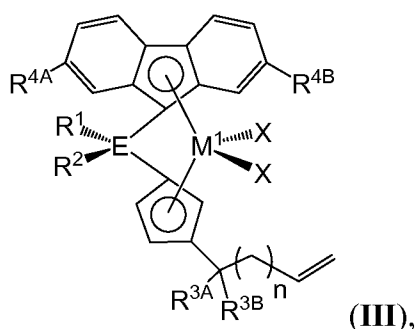
R¹ and R² are independently an aryl group, either of which having up to 10 carbon atoms;

R³ᴬ and R³ᴮ are independently H, methyl, allyl, benzyl, butyl, pentyl, hexyl, or trimethylsilyl;

n is an integer from 1 to 6, inclusive; and

R⁴ᴬ and R⁴ᴮ are independently a hydrocarbonyl group having up to 6 carbon atoms, or hydrogen.

[0071] The at least one *ansa*-metallocene can also comprise, or can be selected from, a compound having the formula:



wherein

M¹ is zirconium or hafnium;

X is Cl, Br, or I;

E is C;

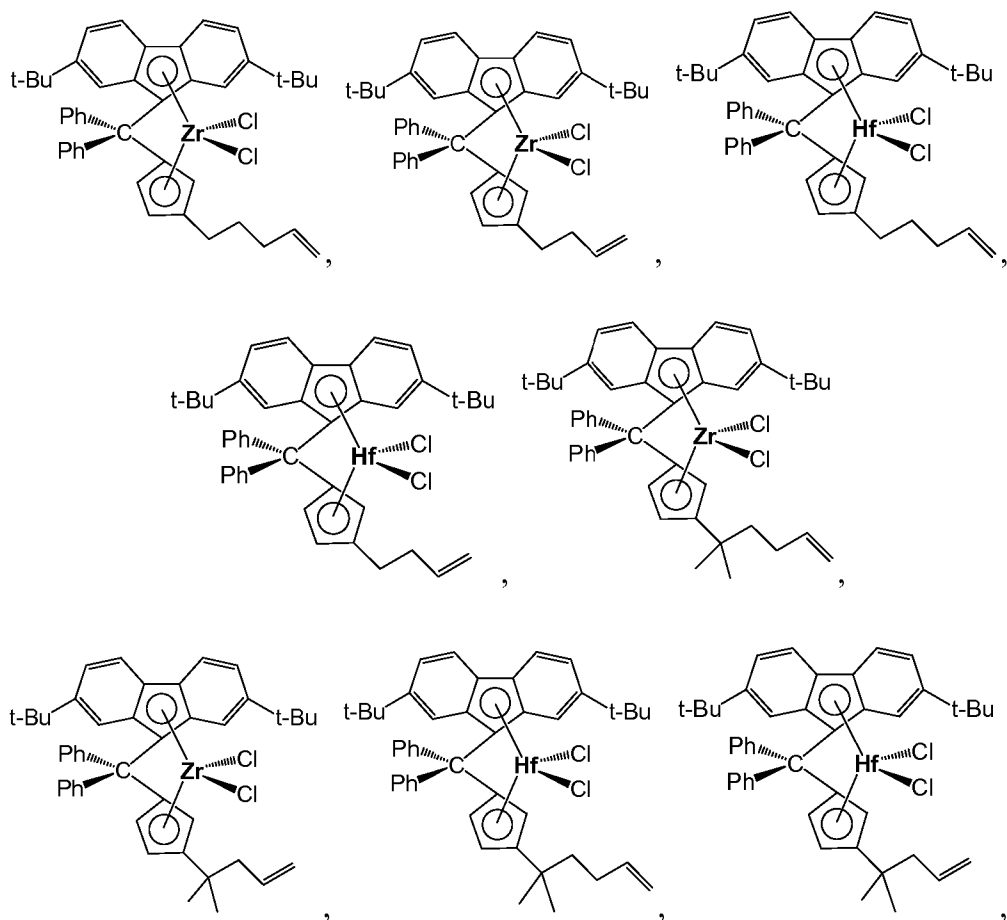
R¹ and R² are independently phenyl;

R³ᴬ and R³ᴮ are independently H or methyl;

n is 1 or 2; and

R⁴ᴬ and R⁴ᴮ are independently H or t-butyl.

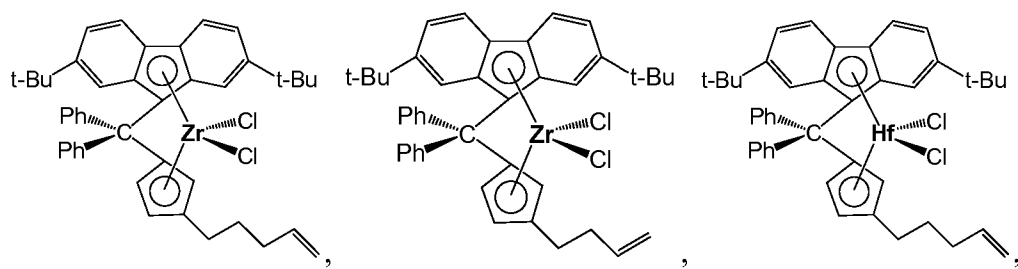
[0072] The at least one *ansa*-metallocene can also comprise, or can be selected from,

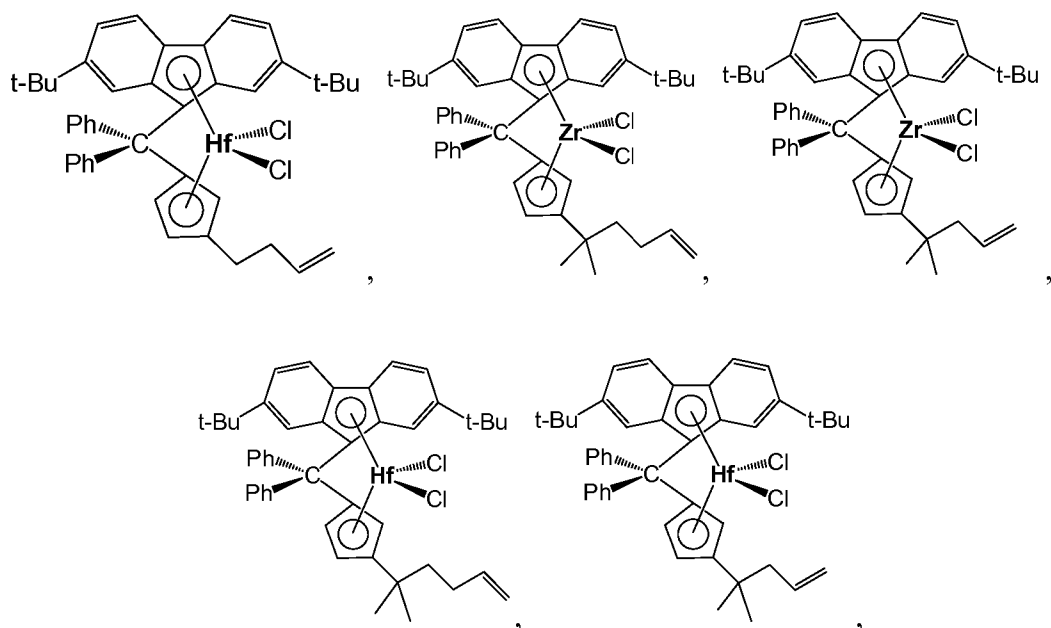


or any combination thereof.

[0073] The catalyst composition can comprise, or the catalyst composition can comprise the contact product of, at least one *ansa*-metallocene, at least one organoaluminum compound, and at least one activator-support, wherein:

a) the at least one *ansa*-metallocene comprises





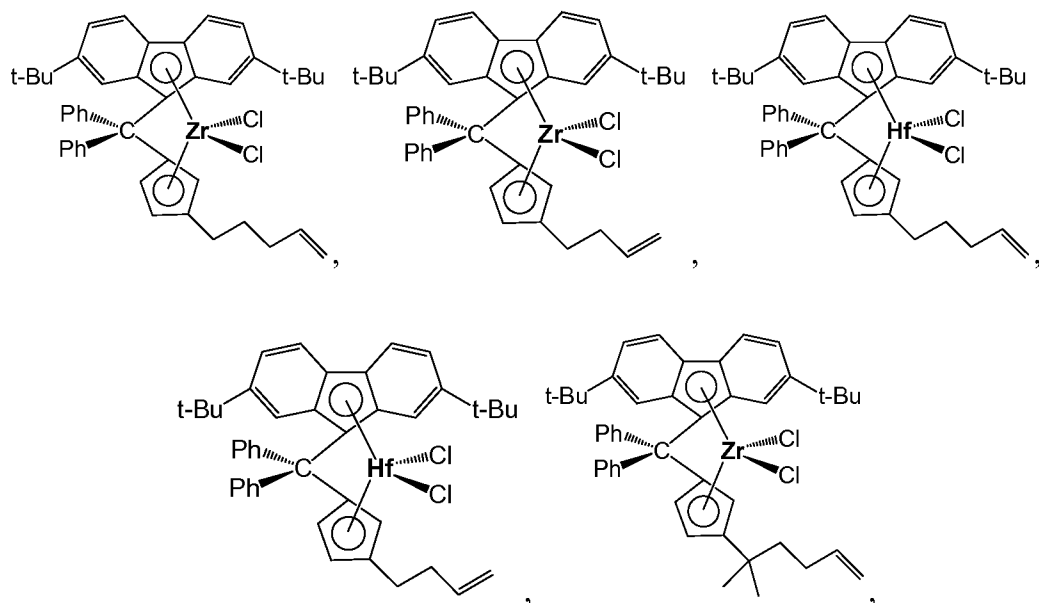
or any combination thereof;

b) the at least one organoaluminum compound comprises triethylaluminum, tri-n-butylaluminum, triisobutylaluminum, or any combination thereof; and

c) the at least one activator-support comprises a sulfated solid oxide.

[0074] The catalyst composition can comprise, or the catalyst composition can comprise the contact product of, at least one *ansa*-metallocene, at least one organoaluminum compound, and at least one activator-support, wherein:

a) the at least one *ansa*-metallocene comprises



or any combination thereof;

b) the at least one organoaluminum compound comprises triethylaluminum, tri-n-butylaluminum, triisobutylaluminum, or any combination thereof; and

c) the at least one activator-support comprises sulfated alumina.

[0075] The catalyst composition can comprise, or the catalyst composition can comprise the contact product of, at least one precontacted *ansa*-metallocene as defined herein, at least one precontacted organoaluminum compound, at least one precontacted olefin, and at least one postcontacted activator-support, wherein each of the *ansa*-metallocene, the organoaluminum compound, the olefin, and the activator-support are as disclosed herein.

The active catalyst composition can be formed typically upon activating the metallocene, which can include converting the metallocene compound to its cationic form, and by providing it with a hydrocarbyl ligand either before, after, or during its conversion to a cation that can initiate olefin polymerization. The at least one reagent that can convert the metallocene into an active catalyst typically includes a component that provides an activatable ligand such as an alkyl to the metallocene, when the metallocene compound does not already comprise such a ligand, and an activator component, as provided herein. In some instances, both functions can be achieved with one component, for example, an organoaluminum compound that can provide an activatable alkyl ligand to the metallocene, and another component that can provide the activator function.

[0076] The activator and alkylation agent for the *ansa*-metallocene compound can be at least one organoaluminum compound, such as methylaluminum compound or isobutylaluminum compound. The activator can be a Lewis acidic organoboron compound capable of abstracting an anionic ligand from the metallocene, such as tris(pentafluorophenyl)boron or triphenylcarbenium tetrakis(pentafluorophenyl)borate, that is used typically in combination with an alkylating agent such as an organoaluminum compound. In still another aspect, a dialkylated tightly-bridged *ansa*-metallocene compound as disclosed herein can be reacted with a Brønsted acidic borate activator such as tri(n-butyl)ammonium tetrakis(p-tolyl)borate or N,N-dimethylanilinium tetrakis(pentafluorophenyl)borate to remove one alkyl ligand to form an alkylated metallocene cation. Yet another aspect provides for a dialkylated tightly-bridged *ansa*-metallocene compound that can be reacted with a Lewis acidic borate activator such as triphenylcarbenium tetrakis(pentafluorophenyl)borate to remove one alkyl ligand to form an alkylated metallocene cation. Thus, while not intending to be bound by theory, it is believed that the active catalyst comprises an alkylated metallocene cation, and any number of reaction steps can be used to generate such a catalyst.

[0077] There is also disclosed a catalyst composition comprising the contact product of at least one tightly-bridged *ansa*-metallocene comprising at least one hydrocarbyl ligand that can initiate olefin polymerization and at least one solid oxide activator-support, without the need for an organoaluminum compound to form the contact product. The *ansa*-metallocene compound is as defined herein and comprises a pendant olefin-containing moiety attached to at least one of the cyclopentadienyl-type ligands, two aryl groups bonded to the bridging atom of the bridging ligand, and at least one hydrocarbyl ligand that can initiate olefin polymerization. An organoaluminum compound would not be required to alkylate this type of "pre-alkylated" *ansa*-metallocene because it already comprises a hydrocarbyl ligand that can initiate olefin polymerization.

Organoaluminum Activators

[0078] There is disclosed a catalyst composition comprising, or a catalyst composition comprising the contact product of, at least one *ansa*-metallocene as defined herein; optionally, at least one organoaluminum compound; and at least one activator, wherein the activator can be selected independently from:

- i) an activator-support comprising a solid oxide treated with an electron-withdrawing anion, a layered mineral, an ion-exchangeable activator-support, or any combination thereof;
- ii) at least one organoaluminum compound;
- iii) at least one organoboron or organoborate compound; or
- iv) any combination thereof.

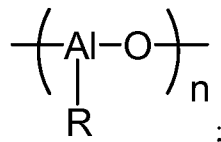
[0079] There is also disclosed a catalyst composition comprising the contact product of at least one *ansa*-metallocene as defined herein; at least one organoaluminum compound; at least one activator-support comprising a solid oxide treated with an electron-withdrawing anion; and optionally, an aluminosilicate cocatalyst. There is also disclosed a catalyst composition comprising an *ansa*-metallocene compound as defined herein containing a pendant unsaturated moiety, an aluminosilicate cocatalyst, an optional activator-support, and an optional organoaluminum compound. The catalyst composition may be substantially free of aluminosilicates, and the catalyst composition may have a polymerization activity in the substantial absence of aluminosilicates.

[0080] There is also disclosed a catalyst composition comprising at least one *ansa*-metallocene compound as described herein and an aluminosilicate. The catalyst composition may not be required to comprise either an acidic activator-support wherein the activator-support comprises a chemically-treated solid oxide, and the catalyst composition is also not required to comprise an organoaluminum compound. Thus, any *ansa*-metallocene compounds disclosed herein can be combined with any of the aluminosilicates (poly(hydrocarbyl aluminum oxides)) disclosed herein, or any combination of aluminosilicates disclosed herein, to form a catalyst composition. Further, any *ansa*-metallocene compounds disclosed herein can be

combined with any aluminoxane or combination of aluminoxanes, and optionally an activator-support; optionally, a layered mineral; optionally, an ion-exchangeable activator-support; optionally, at least one organoboron compound; and optionally, at least one organoborate compound, to form a catalyst composition.

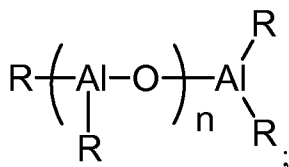
[0081] Aluminoxanes are also referred to as poly(hydrocarbyl aluminum oxides) or organoaluminoxanes. The other catalyst components are typically contacted with the aluminoxane in a saturated hydrocarbon compound solvent, though any solvent which is substantially inert to the reactants, intermediates, and products of the activation step can be used. The catalyst composition formed in this manner can be collected by methods known to those of skill in the art, including but not limited to filtration, or the catalyst composition can be introduced into the polymerization reactor without being isolated.

[0082] The aluminoxane compound may be an oligomeric aluminum compound, wherein the aluminoxane compound can comprise linear structures, cyclic, or cage structures, or typically mixtures of all three. Cyclic aluminoxane compounds having the formula:



wherein

R is a linear or branched alkyl having from 1 to 10 carbon atoms, and n is an integer from 3 to about 10 are encompassed. The $(\text{AlO})_n$ moiety shown here also constitutes the repeating unit in a linear aluminoxane. Thus, linear aluminoxanes having the formula:



wherein

R is a linear or branched alkyl having from 1 to 10 carbon atoms, and n is an integer from 1 to 50, are also encompassed.

[0083] Further, aluminoxanes can also have cage structures of the formula $\text{R}^t_{5m+\alpha} \text{R}^b_{m-\alpha} \text{Al}_{4m} \text{O}_{3m}$, wherein m is 3 or 4 and α is $= n_{\text{Al}(3)} - n_{\text{O}(2)} + n_{\text{O}(4)}$; wherein $n_{\text{Al}(3)}$ is the number of three coordinate aluminum atoms, $n_{\text{O}(2)}$ is the number of two coordinate oxygen atoms, $n_{\text{O}(4)}$ is the number of 4 coordinate oxygen atoms, R^t represents a terminal alkyl group, and R^b represents a bridging alkyl group; wherein R is a linear or branched alkyl having from 1 to 10 carbon atoms.

[0084] Thus, aluminoxanes can be represented generally by formulas such as $(\text{R-Al-O})_n$, $\text{R}(\text{R-Al-O})_n \text{AlR}_2$, wherein the R group is typically a linear or branched $\text{C}_1\text{-C}_6$ alkyl such as methyl, ethyl, propyl, butyl, pentyl, or hexyl wherein n typically represents an integer from 1 to 50. In one embodiment, the aluminoxane compounds include, methylaluminoxane, ethylaluminoxane, n-propylaluminoxane, iso-propylaluminoxane, n-butylaluminoxane, t-butylaluminoxane, sec-butylaluminoxane, isobutylaluminoxane, 1-pentylaluminoxane, 2-pentylaluminoxane, 3-pentylaluminoxane, iso-pentylaluminoxane, neopentylaluminoxane, or combinations thereof.

[0085] While organoaluminoxanes with different types of R groups are encompassed methyl aluminoxane (MAO), ethyl aluminoxane, or isobutyl aluminoxane are typical optional cocatalysts used in the catalyst compositions of this invention. These aluminoxanes are prepared from trimethylaluminum, triethylaluminum, or triisobutylaluminum, respectively, and are sometimes referred to as poly(methyl aluminum oxide), poly(ethyl aluminum oxide), and poly(isobutyl aluminum oxide), respectively. It is also within the scope of the invention to use an aluminoxane in combination with a trialkylaluminum, such as disclosed in U.S. Patent No. 4,794,096.

[0086] The present invention contemplates many values of n in the aluminoxane formulas $(\text{R-Al-O})_n$ and $\text{R}(\text{R-Al-O})_n \text{AlR}_2$, and typically n is at least about 3. However, depending upon how the organoaluminoxane is prepared, stored, and used, the value of n can be variable within a single sample of aluminoxane, and such a combination of organoaluminoxanes are comprised in the methods and compositions of the present invention.

[0087] In preparing the catalyst composition comprising an optional aluminoxane, the molar ratio of the aluminum in the aluminoxane to the metallocene in the composition is usually from 1:10 to 100,000:1. The molar ratio of the aluminum in the aluminoxane to the metallocene in the composition is usually from 5:1 to 15,000:1. The amount of optional aluminoxane added to a polymerization zone is an amount within a range of about 0.01 mg/L to 1000 mg/L, from 0.1 mg/L to 100 mg/L, or from 1 mg/L to 50 mg/L.

[0088] Organoaluminoxanes can be prepared by various procedures which are well known in the art. Examples of organoaluminoxane preparations are disclosed in U.S. Patent Nos. 3,242,099 and 4,808,561. One example of how an aluminoxane can be prepared is as follows. Water, which is dissolved in an inert organic solvent, can be reacted with an aluminum alkyl compound such as AlR_3 to form the desired organoaluminoxane compound. While not intending to be bound by this statement, it is believed that this synthetic method can afford a mixture of both linear and cyclic $(\text{R-Al-O})_n$ aluminoxane species, both of which are encompassed by this invention. Alternatively, organoaluminoxanes can be prepared by reacting an aluminum alkyl compound such as AlR_3 with a hydrated salt, such as hydrated copper sulfate, in an inert organic solvent.

Organoboron and Organoborate Activators

[0089] As provided herein, there is disclosed a catalyst composition comprising, or a catalyst composition comprising the contact product of, at least one *ansa*-metallocene as defined herein; optionally, at least one organoaluminum compound; and at least one activator. The activator can be selected independently from: at least one activator-support as provided herein; at least one organoaluminoxane compound; at least one organoboron or organoborate compound; or any combination thereof. Accordingly, the at least one activator can be selected from at least one organoboron compound, at least one organoborate compound, or a combination thereof.

[0090] There is also disclosed a catalyst composition comprising the contact product of at least one *ansa*-metallocene; at least one organoaluminum compound; at least one activator-support comprising a solid oxide treated with an electron-withdrawing anion; and optionally, an organoboron or organoborate cocatalyst. There is also disclosed a catalyst composition comprising the contact product of: at least one *ansa*-metallocene compound according to the invention containing a pendant unsaturated moiety; an organoboron or organoborate cocatalyst; an organoaluminum compound; and optionally, an activator-support. The catalyst composition is then not required to comprise an activator-support. Any *ansa*-metallocene compound disclosed herein can be combined with any of the organoboron or organoborate cocatalysts disclosed herein, or any combination of organoboron or organoborate cocatalysts disclosed herein, along with a component that provides an activatable ligand such as an alkyl or hydride ligand to the metallocene, when the metallocene compound does not already comprise such a ligand, such as an organoaluminum compound; to form a catalyst composition. Further, any *ansa*-metallocene compounds disclosed herein can be combined with any an organoboron or organoborate cocatalyst; an organoaluminum compound; optionally, at least one aluminoxane; and optionally, an activator-support; to form a catalyst composition of this invention. However, the catalyst composition of this invention may be substantially free of organoboron or organoborate compounds, and in another aspect, the catalyst composition of this invention have polymerization activity in the substantial absence of organoboron or organoborate compounds.

[0091] The term "organoboron" compound can be used to refer to neutral boron compounds, borate salts, or combinations thereof. For example, the organoboron compounds of this invention can comprise a fluoroorgano boron compound, a fluoroorgano borate compound, or a combination thereof. Any fluoroorgano boron or fluoroorgano borate compound known in the art can be utilized. The term fluoroorgano boron compounds has its usual meaning to refer to neutral compounds of the form BY_3 . The term fluoroorgano borate compound also has its usual meaning to refer to the monoanionic salts of a fluoroorgano boron compound of the form $[\text{cation}]^+[\text{BY}_4]^-$, where Y represents a fluorinated organic group. For convenience, fluoroorgano boron and fluoroorgano borate compounds are typically referred to collectively by organoboron compounds, or by either name as the context requires.

[0092] Examples of fluoroorgano borate compounds that can be used as cocatalysts in the present invention include fluorinated aryl borates such as, *N,N*-dimethylanilinium tetrakis(pentafluorophenyl)borate, triphenylcarbenium tetrakis(pentafluorophenyl)borate, lithium tetrakis(pentafluorophenyl)borate, *N,N*-dimethylanilinium tetrakis[3,5-bis(trifluoromethyl)phenyl]borate, triphenylcarbenium tetrakis[3,5-bis(trifluoromethyl)phenyl]borate, including mixtures thereof. Examples of fluoroorgano boron compounds that can be used as cocatalysts in the present invention include tris(pentafluorophenyl)boron, and tris[3,5-bis(trifluoromethyl)-phenyl]boron including mixtures thereof.

[0093] Although not intending to be bound by the following theory, these examples of fluoroorgano borate and fluoroorgano boron compounds, and related compounds, are thought to form "weakly-coordinating" anions when combined with organometal compounds, as disclosed in U.S. Patent 5,919,983.

[0094] Generally, any amount of organoboron compound can be utilized in this invention. The molar ratio of the organoboron compound to the metallocene compound in the composition can be from 0.1:1 to 10:1. Typically, the amount of the fluoroorgano boron or fluoroorgano borate compound used as a cocatalyst for the metallocene is in a range of from 0.5 mole to 10 moles of boron compound per mole of metallocene compound. The amount of fluoroorgano boron or fluoroorgano borate compound used as a cocatalyst for the metallocene can be in a range of from 0.8 mole to 5 moles of boron compound per mole of metallocene compound.

The Optional Ionizing Ionic Compound Cocatalyst

[0095] There is disclosed a catalyst composition comprising, or a catalyst composition comprising the contact product of: 1) at least one tightly-bridged *ansa*-metallocene compound according to the invention comprising an olefin-containing moiety bonded to a cyclopentadienyl-type ligand and at least one aryl group bonded to the bridging atom of the bridging ligand; 2) optionally, at least one organoaluminum compound; and 3) at least one activator, as disclosed herein. There is also described a catalyst composition as disclosed herein, comprising an optional ionizing ionic compound cocatalyst in addition to these other components. The catalyst composition of this invention may be substantially free of ionizing ionic compounds, and the catalyst composition may have polymerization activity in the substantial absence of ionizing ionic compounds. There is disclosed a catalyst composition comprising at least one *ansa*-metallocene compound as disclosed herein, at least one ionizing ionic compound cocatalyst, optionally at least one activator-support, and optionally at least one organoaluminum compound. Examples of ionizing ionic compound are disclosed in U.S. Patent Numbers 5,576,259 and 5,807,938.

[0096] An ionizing ionic compound is an ionic compound which can function to enhance activity of the catalyst composition. While not bound by theory, it is believed that the ionizing ionic compound can be capable of reacting with the metallocene compound and converting the metallocene into a cationic metallocene compound. Again, while not intending to be bound by theory, it is believed that the ionizing ionic compound can function as an ionizing compound by completely or partially extracting an anionic ligand, possibly a non- η^5 -alkadienyl ligand such as (X^3) or (X^4), from the metallocene. However, the ionizing ionic compound is an activator regardless of whether it is ionizes the metallocene, abstracts an (X^3) or (X^4) ligand in a fashion as to form an ion pair, weakens the metal- (X^3) or metal- (X^4) bond in the metallocene, simply coordinates to an (X^3) or (X^4) ligand, or any other mechanisms by which activation can occur. Further, it is not necessary that the ionizing ionic compound activate the metallocene only. The activation function of the ionizing ionic compound is evident in the enhanced activity of catalyst composition as a whole, as compared to a catalyst composition containing catalyst composition that does not comprise any ionizing ionic compound.

[0097] Examples of ionizing ionic compounds include, but are not limited to, the following compounds: tri(n-butyl)ammonium tetrakis(p-tolyl)borate, tri(n-butyl)-ammonium tetrakis(m-tolyl)borate, tri(n-butyl)ammonium tetrakis(2,4-dimethylphenyl)borate, tri(n-butyl)ammonium tetrakis(3,5-dimethylphenyl)borate, tri(n-butyl)ammonium tetrakis[3,5-bis(trifluoromethyl)phenyl]borate, tri(n-butyl)ammonium tetrakis(pentafluorophenyl)borate, N,N-dimethylanilinium tetrakis(p-tolyl)borate, N,N-dimethylanilinium tetrakis(m-tolyl)borate, N,N-dimethylanilinium tetrakis(2,4-dimethylphenyl)borate, N,N-dimethylanilinium tetrakis(3,5-dimethylphenyl)borate, N,N-dimethylanilinium tetrakis[3,5-bis(trifluoromethyl)phenyl]borate, N,N-dimethylanilinium tetrakis(pentafluorophenyl)borate, triphenylcarbenium tetrakis(p-tolyl)borate, triphenylcarbenium tetrakis(m-tolyl)borate, triphenylcarbenium tetrakis(2,4-dimethylphenyl)borate, triphenylcarbenium tetrakis(3,5-dimethylphenyl)borate, triphenylcarbenium tetrakis[3,5-bis(trifluoromethyl)phenyl]borate, triphenylcarbenium tetrakis(pentafluorophenyl)borate, tropylium tetrakis(p-tolyl)borate, tropylium tetrakis(m-tolyl)borate, tropylium tetrakis(2,4-dimethylphenyl)borate, tropylium tetrakis(3,5-dimethylphenyl)borate, tropylium tetrakis[3,5-bis(trifluoromethyl)phenyl]borate, tropylium tetrakis(pentafluorophenyl)borate, lithium tetrakis(pentafluorophenyl)borate, lithium tetrakis(phenyl)borate, lithium tetrakis(p-tolyl)borate, lithium tetrakis(m-tolyl)borate, lithium tetrakis(2,4-dimethylphenyl)borate, lithium tetrakis(3,5-dimethylphenyl)borate, lithium tetrafluoroborate, sodium tetrakis(pentafluorophenyl)borate, sodium tetrakis(phenyl) borate, sodium tetrakis(p-tolyl)borate, sodium tetrakis(m-tolyl)borate, sodium tetrakis(2,4-dimethylphenyl)borate, sodium tetrakis(3,5-dimethylphenyl)borate, sodium tetrafluoroborate, potassium tetrakis(pentafluorophenyl)borate, potassium tetrakis(phenyl)borate, potassium tetrakis(p-tolyl)borate, potassium tetrakis(m-tolyl)borate, potassium tetrakis(2,4-dimethylphenyl)borate, potassium tetrakis(3,5-dimethylphenyl)borate, potassium tetrafluoroborate, triphenylcarbenium tetrakis(p-tolyl)aluminate, triphenylcarbenium tetrakis(m-tolyl)aluminate, triphenylcarbenium tetrakis(2,4-dimethylphenyl)aluminate, triphenylcarbenium tetrakis(3,5-dimethylphenyl)aluminate, triphenylcarbenium tetrakis(pentafluorophenyl)aluminate, tropylium tetrakis(p-tolyl)aluminate, tropylium tetrakis(m-tolyl)aluminate, tropylium tetrakis(2,4-dimethylphenyl)aluminate, tropylium tetrakis(3,5-dimethylphenyl)aluminate, tropylium tetrakis(pentafluorophenyl)aluminate, lithium tetrakis(pentafluorophenyl)aluminate, lithium tetrakis(phenyl)aluminate, lithium tetrakis(p-tolyl)aluminate, lithium tetrakis(m-tolyl)aluminate, lithium tetrakis(2,4-dimethylphenyl)aluminate, lithium tetrakis(3,5-dimethylphenyl)aluminate, lithium tetrafluoroaluminate, sodium tetrakis(pentafluorophenyl)aluminate, sodium tetrakis(phenyl)aluminate, sodium tetrakis(p-tolyl)aluminate, sodium tetrakis(m-tolyl)aluminate, sodium tetrakis(2,4-dimethylphenyl)aluminate, sodium tetrakis(3,5-dimethylphenyl)aluminate, sodium tetrafluoroaluminate, potassium tetrakis(pentafluorophenyl)aluminate, potassium tetrakis(phenyl)aluminate, potassium tetrakis(p-tolyl)aluminate, potassium tetrakis(m-tolyl)aluminate, potassium tetrakis(2,4-dimethylphenyl)aluminate, potassium tetrakis(3,5-dimethylphenyl)aluminate, potassium tetrafluoroaluminate, triphenylcarbenium tris(2,2',2"-nonafluorobiphenyl)fluoroaluminate, silver tetrakis(1,1,1,3,3,3-hexafluoroisopropanolato)aluminate, or silver tetrakis(perfluoro-t-butoxy)aluminate, or any combination thereof. However, these ionizing ionic compound are exemplary, and the ionizing ionic compound is not limited thereto in the present invention.

The Olefin Monomer

[0098] Unsaturated reactants that are useful in the polymerization processes with catalyst compositions and processes disclosed herein include olefin compounds having from 2 to 30 carbon atoms per molecule and having at least one olefinic double bond. There are disclosed homopolymerization processes using a single olefin such as ethylene or propylene, as well as copolymerization reactions with at least one different olefinic compound. A copolymerization reaction of ethylene, copolymers of ethylene comprise a major amount of ethylene (>50 mole percent) and a minor amount of comonomer (<50 mole percent), though this is not a requirement. The comonomers that can be copolymerized with ethylene should have from three to 20 carbon atoms in their molecular chain.

[0099] Acyclic, cyclic, polycyclic, terminal (α), internal, linear, branched, substituted, unsubstituted, functionalized, and non-functionalized olefins can be employed. For example, typical unsaturated compounds that can be polymerized with the catalysts disclosed herein include, but are not limited to, propylene, 1-butene, 2-butene, 3-methyl-1-butene, isobutylene, 1-pentene, 2-pentene, 3-methyl-1-pentene, 4-methyl-1-pentene, 1-hexene, 2-hexene, 3-hexene, 3-ethyl-1-hexene, 1-heptene, 2-heptene, 3-heptene, the four normal octenes, the four normal nonenes, the five normal decenes, and mixtures of any two or more thereof. Cyclic and bicyclic olefins, including cyclopentene, cyclohexene, norbornylene, and norbornadiene, can also be polymerized as described above.

[0100] When a copolymer is desired, the monomer ethylene can be copolymerized with a comonomer. Examples of the comonomer include propylene, 1-butene, 2-butene, 3-methyl-1-butene, isobutylene, 1-pentene, 2-pentene, 3-methyl-1-pentene, 4-methyl-1-pentene, 1-hexene, 2-hexene, 3-hexene, 3-ethyl-1-hexene, 1-heptene, 2-heptene, 3-heptene, the four normal octenes, the four normal nonenes, or the five normal decenes. The comonomer can be 1-butene, 1-pentene, 1-hexene, 1-octene, 1-decene, or styrene.

[0101] The amount of comonomer introduced into a reactor zone to produce the copolymer may generally be from 0.001 to 99 weight percent comonomer based on the total weight of the monomer and comonomer. The amount of comonomer introduced into a reactor zone to produce the copolymer may generally be from 0.01 to 50 weight percent comonomer based on the total weight of the monomer and comonomer. The amount of comonomer introduced into a reactor zone may be from 0.01 to 10 weight percent comonomer, and from 0.1 to 5 weight percent comonomer based on the total weight of the monomer and comonomer. Alternatively, an amount sufficient to give the above described concentrations by weight, in the copolymer produced can be used.

[0102] While not intending to be bound by this theory, in the event that branched, substituted, or functionalized olefins are used as reactants, it is believed that steric hindrance can impede and/or slow the polymerization process. Thus, branched and/or cyclic portion(s) of the olefin removed somewhat from the carbon-carbon double bond would not be expected to hinder the reaction in the way that the same olefin substituents situated more proximate to the carbon-carbon double bond might. At least one reactant for the catalyst compositions of this invention may be ethylene, so the polymerizations are either homopolymerizations or copolymerizations with a different acyclic, cyclic, terminal, internal, linear, branched, substituted, or unsubstituted olefin. In addition, the catalyst compositions of this invention can be used in polymerization of diolefin compounds, including but are not limited to, 1,3-butadiene, isoprene, 1,4-pentadiene, and 1,5-hexadiene.

Preparation of the Catalyst Composition

[0103] There is disclosed a catalyst composition and a method comprising contacting at least one tightly-bridged *ansa*-metallocene compound, at least one activator, and optionally at least one organoaluminum compound, as disclosed herein. The method disclosed herein encompasses any series of contacting steps that allows contacting each of the components provided, any order of contacting components or mixtures of components. While not intending to be limiting, examples of contacting steps are typically exemplified herein using a treated solid oxide activator-support and an organoaluminum cocatalyst. These exemplary steps can encompass any number of precontacting and postcontacting steps, and can further encompass using an olefin monomer as a contact component in any of these steps. Examples of preparative methods to prepare the catalyst composition of this invention follow.

[0104] The *ansa*-metallocene can be precontacted with an olefinic monomer, not necessarily the olefin monomer to be polymerized, and an organoaluminum cocatalyst for a first period of time prior to contacting this precontacted mixture with the solid oxide activator-support. For example, the first period of time for contact, the precontact time, between the *ansa*-metallocene, the olefinic monomer, and the organoaluminum cocatalyst can typically range from time 1 minute to 24 hours, and from 0.1 to 1 hour is typical. Precontact times from 10 minutes to 30 minutes are also typical.

[0105] Once the precontacted mixture of *ansa*-metallocene, olefin monomer, and organoaluminum cocatalyst is contacted with the solid oxide activator, this composition (further comprising the solid oxide activator) is termed the postcontacted mixture. Typically, the postcontacted mixture can be allowed to remain in contact for a second period of time, the postcontact time, prior to being initiating the polymerization process. Postcontact times between the solid oxide activator-support and the precontacted mixture typically range from time 1 minute to 24 hours, and from 0.1 to 1 hour

is typical. Postcontact times from 10 minutes to 30 minutes are also typical.

[0106] The various catalyst components (for example, *ansa*-metallocene, activator-support, organoaluminum cocatalyst, and optionally an unsaturated hydrocarbon) can be contacted in the polymerization reactor simultaneously while the polymerization reaction is proceeding. Alternatively, any two or more of these catalyst components can be "precontacted" in a vessel or tube prior to their entering the reaction zone. This precontacting step can be a continuous process, in which the precontacted product is fed continuously to the reactor, or it can be a stepwise or batchwise process in which a batch of precontacted product can be added to make a catalyst composition. This precontacting step can be carried out over a time period that can range from a few seconds to as much as several days, or longer. The continuous precontacting step can last typically from 1 second to 1 hour. Also, the continuous precontacting step can last typically from 10 seconds to 45 minutes, or from 1 minute to 30 minutes.

[0107] Alternatively the precontacting process can be carried out in multiple steps, rather than a single step, in which multiple mixtures are prepared, each comprising a different set of catalyst components. For example, at least two catalyst components can be contacted forming a first mixture, followed by contacting the first mixture with at least one other catalyst component forming a second mixture, and so forth.

[0108] Multiple precontacting steps can be carried out in a single vessel or in multiple vessels. Further, multiple precontacting steps can be carried out in series (sequentially), in parallel, or a combination thereof. For example, a first mixture of two catalyst components can be formed in a first vessel, a second mixture comprising the first mixture plus one additional catalyst component can be formed in the first vessel or in a second vessel, which is typically placed downstream of the first vessel.

[0109] One or more of the catalyst components can be split and used in different precontacting treatments. For example, part of a catalyst component can be fed into a first precontacting vessel for precontacting with at least one other catalyst component, while the remainder of that same catalyst component can be fed into a second precontacting vessel for precontacting with at least one other catalyst component, or can be fed directly into the reactor, or a combination thereof. The precontacting can be carried out in any suitable equipment, such as tanks, stirred mix tanks, various static mixing devices, a tube, a flask, a vessel of any type, or any combination thereof.

[0110] For example, a catalyst composition of this invention is prepared by contacting 1-hexene, triisobutylaluminum or tri-*n*-butylaluminum, and an *ansa*-metallocene for at least 30 minutes, followed by contacting this precontacted mixture with a sulfated alumina activator-support for at least 10 minutes up to one hour to form the active catalyst.

[0111] The precontacting step typically increases the productivity of the polymer as compared to the same catalyst composition that is prepared without this precontacting step. The enhanced activity catalyst composition of this invention can be used for homopolymerization of an α -olefin monomer such as ethylene or copolymerization of an α -olefin and a comonomer. However, neither a precontacting step nor a postcontacting step are required for this invention.

[0112] The postcontacted mixture can be heated at a temperature and for a duration sufficient to allow adsorption, impregnation, or interaction of precontacted mixture and the solid oxide activator-support, such that a portion of the components of the precontacted mixture is immobilized, adsorbed, or deposited thereon. For example, the postcontacted mixture can be heated from between -17.7°C to 82.2°C (0°F to 150°F). Temperatures between 4.4°C to 35°C (40°F to 95°F) are typical if the mixture is heated at all.

[0113] The molar ratio of the *ansa*-metallocene compound to the organoaluminum compound can be from 1:1 to 1:10,000. The molar ratio of the *ansa*-metallocene compound to the organoaluminum compound can be from 1:1 to 1:1,000, and from 1:1 to 1:100. These molar ratios reflect the ratio of *ansa*-metallocene compound to the total amount of organoaluminum compound in both the precontacted mixture and the postcontacted mixture combined.

[0114] When a precontacting step is used, generally, the molar ratio of olefin monomer to *ansa*-metallocene compound in the precontacted mixture can be from 1:10 to 100,000:1, or from 10:1 to 1,000:1.

[0115] The weight ratio of the solid oxide activator to the organoaluminum compound can range from 1:5 to 1,000:1. The weight ratio of the solid oxide activator to the organoaluminum compound can be from 1:3 to 100:1, and from 1:1 to 50:1.

[0116] The weight ratio of the *ansa*-metallocene to solid oxide activator-support can be from 1:1 to 1:1,000,000. The weight ratio of the *ansa*-metallocene to solid oxide activator-support can be from 1:10 to 1:100,000, and from 1:20 to 1:1000.

[0117] Aluminoxane is not required to form the catalyst composition disclosed herein, a feature that allows lower polymer production costs. Accordingly, the present invention can use AlR_3 -type organoaluminum compounds and an activator-support in the absence of aluminoxanes. While not intending to be bound by theory, it is believed that the organoaluminum compounds likely do not activate the metallocene catalysts in the same manner as an organoaluminum compound.

[0118] Additionally, no expensive borate compounds or MgCl_2 are required to form the catalyst composition of this invention, although aluminoxane, borate compounds, MgCl_2 , or any combination thereof can be used optionally in the catalyst composition of this invention. Further, cocatalysts such as aluminoxanes, organoboron compounds, ionizing ionic compounds, or any combination thereof can be used as cocatalysts with the *ansa*-metallocene, either in the

presence or absence of the activator support. Further, cocatalysts such as aluminoxanes, organoboron compounds, ionizing ionic compounds, or any combination thereof can be used as with the *ansa*-metallocene, either and in the presence or absence of an organoaluminum compound, as specified herein. Thus, the at least one organoaluminum compound is optional when at least one ligand on the metallocene is a hydrocarbyl group, H, or BH₄; when the at least one activator comprises an organoaluminoxane compound; or both these conditions are present. However, the catalyst compositions of the present invention are active in the substantial absence of cocatalysts such as aluminoxanes, organoboron compounds, ionizing ionic compounds, or any combination thereof.

[0119] There is disclosed a process to produce a catalyst composition, comprising:

contacting at least one *ansa*-metallocene as defined herein, at least one olefin, and at least one organoaluminum compound for a first period of time to form a precontacted mixture comprising at least one precontacted *ansa*-metallocene, at least one precontacted organoaluminum compound, and at least one precontacted olefin; and contacting the precontacted mixture with at least one activator-support and optionally additional organoaluminum compound for a second period of time to form a postcontacted mixture comprising at least one postcontacted *ansa*-metallocene, at least one postcontacted organoaluminum compound, at least one postcontacted olefin, and at least one postcontacted activator-support.

[0120] The catalytic activity of the catalyst of this invention is typically greater than or equal to 100 grams polyethylene per gram of chemically treated solid oxide per hour (abbreviated gP/(gCTSO·hr)). The catalyst of this invention can be characterized by an activity of greater than or equal to 250 gP/(gCTSO·hr), and an activity of greater than or equal to 500 gP/(gCTSO·hr). The catalyst of this invention can be characterized by an activity of greater than or equal to 1000 gP/(gCTSO·hr), and an activity of greater than or equal to 2000 gP/(gCTSO·hr). This activity is typically measured under slurry polymerization conditions using isobutane as the diluent, with a polymerization temperature of 90 °C, and an ethylene pressure of 3792.1 kPa (550 psig). This activity is measured under slurry polymerization conditions using isobutane as the diluent, with a polymerization temperature from 80 °C to 105 °C, and an ethylene pressure from 3102.6 kPa (450 psig) to 3792.1 kPa (550 psig). The reactor should have substantially no indication of any wall scale, coating or other forms of fouling upon making these measurements.

Utility of the Catalyst Composition in Polymerization Processes

[0121] The catalysts of the present invention are intended for any olefin polymerization method known in the art, using various types of polymerization reactors. As used herein, "polymerization reactor" includes any polymerization reactor capable of polymerizing olefin monomers to produce homopolymers or copolymers. Such homopolymers and copolymers are referred to as resins or polymers. The various types of reactors include those that may be referred to as batch, slurry, gas-phase, solution, high pressure, tubular or autoclave reactors. Gas phase reactors may comprise fluidized bed reactors or staged horizontal reactors. Slurry reactors may comprise vertical or horizontal loops. High pressure reactors may comprise autoclave or tubular reactors. Reactor types can include batch or continuous processes. Continuous processes could use intermittent or continuous product discharge. Processes may also include partial or full direct recycle of un-reacted monomer, un-reacted comonomer, and/or diluent.

[0122] Polymerization reactor systems of the present invention may comprise one type of reactor in a system or multiple reactors of the same or different type. Production of polymers in multiple reactors may include several stages in at least two separate polymerization reactors interconnected by a transfer device making it possible to transfer the polymers resulting from the first polymerization reactor into the second reactor. The desired polymerization conditions in one of the reactors may be different from the operating conditions of the other reactors. Alternatively, polymerization in multiple reactors may include the manual transfer of polymer from one reactor to subsequent reactors for continued polymerization. Multiple reactor systems may include any combination including, but not limited to, multiple loop reactors, multiple gas reactors, a combination of loop and gas reactors, multiple high pressure reactors or a combination of high pressure with loop and/or gas reactors. The multiple reactors may be operated in series or in parallel.

[0123] The polymerization reactor system may comprise at least one loop slurry reactor. Such reactors are known in the art and may comprise vertical or horizontal loops. Monomer, diluent, catalyst and optionally any comonomer may be continuously fed to a loop reactor where polymerization occurs. Generally, continuous processes may comprise the continuous introduction of a monomer, a catalyst, and a diluent into a polymerization reactor and the continuous removal from this reactor of a suspension comprising polymer particles and the diluent. Reactor effluent may be flashed to remove the solid polymer from the liquids that comprise the diluent, monomer and/or comonomer. Various technologies may be used for this separation step including but not limited to, flashing that may include any combination of heat addition and pressure reduction; separation by cyclonic action in either a cyclone or hydrocyclone; or separation by centrifugation.

[0124] A typical slurry polymerization process (also known as the particle form process), which is well known in the art is disclosed, for example, in U.S. Patent Nos. 3,248,179, 4,501,885, 5,565,175, 5,575,979, 6,239,235, 6,262,191

and 6,833,415.

[0125] Suitable diluents used in slurry polymerization are well known in the art and include, but are not limited to, the monomer being polymerized and hydrocarbons that are liquids under reaction conditions. Examples of suitable diluents include, but are not limited to, hydrocarbons such as propane, cyclohexane, isobutane, n-butane, n-pentane, isopentane, neopentane, and n-hexane. Some loop polymerization reactions can occur under bulk conditions where no diluent is used. An example is polymerization of propylene monomer as disclosed in U.S. Patent Nos. 5,455,314.

[0126] The polymerization reactor may comprise at least one gas phase reactor. Such systems are known in the art and may employ a continuous recycle stream containing one or more monomers continuously cycled through a fluidized bed in the presence of the catalyst under polymerization conditions. A recycle stream may be withdrawn from the fluidized bed and recycled back into the reactor. Simultaneously, polymer product may be withdrawn from the reactor and new or fresh monomer may be added to replace the polymerized monomer. Such gas phase reactors may comprise a process for multi-step gas-phase polymerization of olefins, in which olefins are polymerized in the gaseous phase in at least two independent gas-phase polymerization zones while feeding a catalyst-containing polymer formed in a first polymerization zone to a second polymerization zone. One type of gas phase reactor is disclosed in U.S. Patent Nos. 5,352,749, 4588,790 and 5,436,304.

[0127] A high pressure polymerization reactor may comprise a tubular reactor or an autoclave reactor, both of which are known in the art. Tubular reactors may have several zones where fresh monomer, initiators, or catalysts are added. Monomer may be entrained in an inert gaseous stream and introduced at one zone of the reactor. Initiators, catalysts, and/or catalyst components may be entrained in a gaseous stream and introduced at another zone of the reactor. The gas streams may be intermixed for polymerization. Heat and pressure may be employed appropriately to obtain optimal polymerization reaction conditions.

[0128] The polymerization reactor may comprise a solution polymerization reactor wherein the monomer is contacted with the catalyst composition by suitable stirring or other means. A carrier comprising an inert organic diluent or excess monomer may be employed. If desired, the monomer may be brought in the vapor phase into contact with the catalytic reaction product, in the presence or absence of liquid material. The polymerization zone is maintained at temperatures and pressures that will result in the formation of a solution of the polymer in a reaction medium. Agitation may be employed to obtain better temperature control and to maintain uniform polymerization mixtures throughout the polymerization zone. Adequate means are utilized for dissipating the exothermic heat of polymerization. These reactors are known in the art.

[0129] Polymerization reactors suitable for the present invention may further comprise any combination of at least one raw material feed system, at least one feed system for catalyst or catalyst components, and/or at least one polymer recovery system. Suitable reactor systems for the present invention may further comprise systems for feedstock purification, catalyst storage and preparation, extrusion, reactor cooling, polymer recovery, fractionation, recycle, storage, loadout, laboratory analysis, and process control.

[0130] Conditions that are controlled for polymerization efficiency and to provide resin properties include temperature, pressure and the concentrations of various reactants. Polymerization temperature can affect catalyst productivity, polymer molecular weight and molecular weight distribution. Suitable polymerization temperature may be any temperature below the de-polymerization temperature according to the Gibbs Free energy equation. Typically this includes from 60°C to 280°C, for example, and from 70°C to 110°C, depending upon the type of polymerization reactor.

[0131] Suitable pressures will also vary according to the reactor and polymerization type. The pressure for liquid phase polymerizations in a loop reactor is typically less than 6894.8 kPa (1000 psig). Pressure for gas phase polymerization is usually at 1379 kPa - 3447.4 kPa (200 - 500 psig). High pressure polymerization in tubular or autoclave reactors is generally run at 137895.1 kPa to 517107 kPa (20,000 to 75,000 psig). Polymerization reactors can also be operated in a supercritical region occurring at generally higher temperatures and pressures. Operation above the critical point of a pressure/temperature diagram (supercritical phase) may offer advantages.

[0132] The concentration of various reactants can be controlled to produce resins with certain physical and mechanical properties. The proposed end-use product that will be formed by the resin and the method of forming that product determines the desired resin properties. Mechanical properties include tensile, flexural, impact, creep, stress relaxation and hardness tests. Physical properties include density, molecular weight, molecular weight distribution, melting temperature, glass transition temperature, temperature melt of crystallization, density, stereoregularity, crack growth, long chain branching and rheological measurements.

[0133] The concentrations of monomer, co-monomer, hydrogen, co-catalyst, modifiers, and electron donors are important in producing these resin properties. Comonomer is used to control product density. Hydrogen is used to control product molecular weight. Co-catalysts can be used to alkylate, scavenge poisons and control molecular weight. Modifiers can be used to control product properties and electron donors affect stereoregularity. In addition, the concentration of poisons must be minimized since they impact the reactions and product properties.

[0134] The polymer or resin may be formed into various articles, including bottles, drums, toys, household containers, utensils, film products, drums, fuel tanks, pipes, geomembranes, and liners. Various processes may be used to form these articles, including blow molding, extrusion molding, rotational molding, thermoforming, and cast molding. After

polymerization, additives and modifiers can be added to the polymer to provide better processing during manufacturing and for desired properties in the end product. Additives include surface modifiers such as slip agents, antiblocks, tackifiers; antioxidants such as primary and secondary antioxidants; pigments; processing aids such as waxes/oils and fluoroelastomers; and special additives such as fire retardants, antistats, scavengers, absorbers, odor enhancers, and degradation agents.

Ethylene Polymers Prepared According to this Invention

[0135] The ethylene polymer produced using the catalyst composition of this invention is typically characterized by lower levels of long chain branching (LCB) than are typically observed when using a tightly-bridged *ansa*-metallocene compound without an olefin-containing moiety bonded to a cyclopentadienyl-type ligand, even when the comparative metallocene comprises at least one aryl group bonded to the bridging atom of the bridging ligand. The ethylene polymer produced using the catalyst composition of this invention is typically characterized by higher molecular weights than are typically observed when using a tightly-bridged *ansa*-metallocene compound without at least one aryl group bonded to the bridging atom of the bridging ligand, even when the comparative metallocene comprises an olefin-containing moiety bonded to a cyclopentadienyl-type ligand. Figures 3 through 8 illustrate various aspects of olefin homopolymer produced according to this invention.

[0136] The combination of size exclusion chromatography (SEC) and multiangle light scattering (MALS) detection was used to detect and characterize polymer branching. As illustrated in Figures 3-5, diagrams obtained from SEC-MALS analysis plotting the radius of gyration (R_g), one measurement of molecular size, versus M_w , of the ethylene homopolymer produced in Examples 1-7 and Examples 10-11, demonstrate one aspect of the invention's utility in reducing LCB. Deviation of the radius of gyration (R_g) from a known linear control (in this instance, HiD9640) indicates branching. Thus, the data from Figures 3-5 demonstrate that the polymers prepared using the catalyst compositions according to this invention are only very slightly deviated from the linear standard, HiD9640, at high molecular weight end in the plot of R_g vs. M_w .

[0137] Figures 6 and 7 illustrate plots of $\log(\eta_0)$ versus $\log(M_w)$ for polymers prepared according to inventive Examples 1-11 and comparative Examples 14-16, respectively, and further illustrate how reduced LCB levels are manifested (see: Table 1). Linear polyethylene polymers are observed to follow a power law relationship between their zero shear viscosity, η_0 , and their weight average molecular weight, M_w , with a power very close to 3.4. This relationship is shown by a straight line with a slope of 3.4 when the logarithm of η_0 is plotted versus the logarithm of M_w . Deviations from this linear polymer line are generally accepted as being caused by the presence of long-chain branching (LCB). Janzen and Colby presented a model that predicts the expected deviation from the linear plot of $\log(\eta_0)$ vs. $\log(M_w)$ for given frequencies of LCB as a function of the weight average molecular weight of the polymer. See: ["Diagnosing long-chain branching in polyethylenes," J. Mol. Struct. 485-486, 569-584 (1999)].

[0138] Thus, Figures 6 and 7 plot the logarithm of η_0 versus the logarithm of M_w for polymers prepared according to this invention, illustrating the dependence of zero shear melt viscosity on weight average molecular weight (M_w), and demonstrating that these polymers are only very slightly deviated from the well-known 3.4 power law "Arnett line" which is used as an indication of a linear polymer (J. Phys. Chem. 1980, 84, 649). Consistent with this observation, both SEC-MALS and rheology data indicate that the inventive metallocenes produce very low LCB in ethylene polymerization, as illustrated in Figure 6 for Examples 1-11. In contrast, polymers made according to the comparative Examples 14-16 had much lower M_w relative to polymers prepared according to this invention, Examples 1-11. Typically, these polymers also had similar or slightly higher levels of LCB, as illustrated in Figure 7.

[0139] Figure 8 demonstrates a comparison of gel-permeation chromatography (GPC) runs for the polymers produced according to Examples 1-11 and Examples 14-16, of this invention. These GPC results (Table 1 and Figure 8) indicate that the polyethylenes (PE) produced according to this invention generally have high molecular weights. While the polymers made according to comparative Examples 14-16 were characterized by low levels of LCB (Figure 8x), these comparative polymers had relatively lower M_w as compared to the polymers prepared according to this invention. Comparative Examples 12 and 13 also demonstrate that the catalyst prepared using comparative metallocene C-1 exhibited poor activity (Table 1). Further, a considerable amount of insoluble polymers (about 50 wt%) were observed in the polymeric samples prepared according to Comparative Examples 12 and 13 during the process of preparing the GPC and SEC-MALS samples of these materials. Thus, using polymeric samples prepared according to Comparative Examples 12 and 13, 25 mg to 28 mg of polymer was mixed in 25 mL of 1,2,4-trichlorobenzene, and agitated for 5 hours, while maintaining this mixture at 150°C. Visual examination of the sample vial containing the sample prepared as described revealed that a precipitate formed on the sides of the sample vial. This observation indicates that the polymers prepared according to Examples 12 and 13, using metallocene C-1, were non-linear polymers. No insoluble polymers were observed in polymers prepared according to any other Example.

Definitions

[0140] In order to more clearly define the terms used herein, the following definitions are provided. To the extent that any definition or usage provided by any document incorporated herein by reference conflicts with the definition or usage provided herein, the definition or usage provided herein controls.

[0141] The term "polymer" is used herein to mean homopolymers comprising ethylene and/or copolymers of ethylene and another olefinic comonomer. "Polymer" is also used herein to mean homopolymers and copolymers of any other polymerizable monomer disclosed herein.

[0142] The term "cocatalyst" is generally used herein to refer to the organoaluminum compounds that can constitute one component of the catalyst composition, but also refers to the optional components of the catalyst composition including, but not limited to, aluminoxanes, organoboron compounds, organoborate compounds, or ionizing ionic compounds, as disclosed herein. In one aspect, cocatalysts can be organoaluminum compounds of the formula $\text{Al}(\text{X}^5)_n(\text{X}^6)_{3-n}$, wherein (X^5) is a hydrocarbyl having from 1 to about 20 carbon atoms; (X^6) is alkoxide or aryloxide, any of which having from 1 to about 20 carbon atoms, halide, or hydride; and n is a number from 1 to 3, inclusive. The term cocatalyst can be used regardless of the actual function of the compound or any chemical mechanism by which the compound might operate.

[0143] The term "precontacted" mixture is used herein to describe a first mixture of catalyst components that are contacted for a first period of time prior to the first mixture being used to form a "postcontacted" or second mixture of catalyst components that are contacted for a second period of time. Typically, the precontacted mixture describes a mixture of metallocene, olefin monomer, and organoaluminum compound, before this mixture is contacted with the acidic activator-support and optionally an organoaluminum compound. Thus, "precontacted" describes components that are used to contact each other, but prior to contacting the components in the second, postcontacted mixture. Accordingly, this invention may occasionally distinguish between a component used to prepare the precontacted mixture and that component after the mixture has been prepared. For example, according to this description, it is possible for the precontacted organoaluminum compound, once it is contacted with the metallocene and the olefin monomer, to have reacted to form at least one different chemical compound, formulation, or structure from the distinct organoaluminum compound used to prepare the precontacted mixture. In this case, the precontacted organoaluminum compound or component is described as comprising an organoaluminum compound that was used to prepare the precontacted mixture.

[0144] Similarly, the term "postcontacted" mixture is used herein to describe a second mixture of catalyst components that are contacted for a second period of time, and one constituent of which is the "precontacted" or first mixture of catalyst components that were contacted for a first period of time. Typically, the term "postcontacted" mixture is used herein to describe the mixture of metallocene, olefin monomer, organoaluminum compound, and acidic activator-support, formed from contacting the precontacted mixture of a portion of these components with the any additional components added to make up the postcontacted mixture. Generally, the additional component added to make up the postcontacted mixture is the solid oxide activator, and optionally can include an organoaluminum compound the same or different from the organoaluminum compound used to prepare the precontacted mixture, as described herein. Accordingly, this invention may also occasionally distinguish between a component used to prepare the postcontacted mixture and that component after the mixture has been prepared.

[0145] The terms "catalyst composition," and "catalyst mixture," do not depend upon the actual product of the reaction of the components of the mixtures, the nature of the active catalytic site, or the fate of the aluminum cocatalyst, *ansa*-metallocene, any olefin monomer used to prepare a precontacted mixture, or the solid oxide activator after combining these components. Therefore, the terms catalyst composition, and catalyst mixture include both heterogeneous compositions and homogenous compositions.

[0146] The term "hydrocarbyl" is used to specify a hydrocarbon radical group that includes aryl, alkyl, cycloalkyl, alkenyl, cycloalkenyl, cycloalkadienyl, alkynyl, aralkyl, aralkenyl, aralkynyl and includes all substituted, unsubstituted, branched, linear, heteroatom substituted derivatives thereof. Unless otherwise specified, the hydrocarbyl groups of this invention typically comprise up to 20 carbon atoms. In one aspect, hydrocarbyl groups can have up to 12 carbon atoms, up to 8 carbon atoms, or up to 6 carbon atoms.

[0147] The term "hydrocarbyloxy" group is used generically to refer collectively to both alkoxide and aryloxide groups. Unless otherwise specified, the hydrocarbyl oxide groups of this invention typically comprise up to 20 carbon atoms. In one aspect, hydrocarbyloxy groups can have up to 12 carbon atoms, up to 8 carbon atoms, or up to 6 carbon atoms.

[0148] The term "hydrocarbylamino" group is used generically to refer collectively to alkylamino (NHR), arylamino (NHR), dialkylamino (NR_2), and diarylamino (NR_2) groups. Unless otherwise specified, the hydrocarbyl amino groups of this invention typically comprise up to 20 carbon atoms. In one aspect, hydrocarbylamino groups can have up to 12 carbon atoms, up to 8 carbon atoms, or up to 6 carbon atoms.

[0149] The term "alkenyl" is used broadly to specify a hydrocarbyl group that comprises an alkene moiety, regardless of the particular regiochemistry of the alkene moiety and encompassing all stereochemical isomers. Thus, for example, the term alkenyl is intended to include any $\text{CH}=\text{CH}_2$ -substituted or $\text{CH}=\text{CMe}_2$ -substituted alkyl group, regardless of

where the substitution occurs within the alkyl group. Terms such as olefin-containing hydrocarbyl group or olefin-containing pendant group are typically used interchangeably with alkenyl group, again illustrating that these terms are not intended to be bound by the particular location of the C=C double bond within the group. Unless otherwise specified, the alkenyl groups of this invention typically comprise up to 20 carbon atoms. In one aspect, alkenyl groups can have up to 12 carbon atoms, up to 8 carbon atoms, or up to 6 carbon atoms.

[0150] The terms solid oxide activator-support, acidic activator-support, activator-support, treated solid oxide, and treated solid oxide compound are used herein to indicate a treated, solid, inorganic oxide of relatively high porosity, which exhibits Lewis acidic or Brønsted acidic behavior, and which has been treated with an electron-withdrawing component, typically an anion, and which is calcined. The electron-withdrawing component is typically an electron-withdrawing anion source compound. Thus, the treated solid oxide compound comprises the calcined contact product of at least one solid oxide compound with at least one electron-withdrawing anion source compound. Typically, the activator-support or "treated solid oxide compound" comprises at least one ionizing, acidic solid oxide compound. The terms support or activator-support are not used to imply these components are inert, and this component should not be construed as an inert component of the catalyst composition.

[0151] The term "activator," as used herein, refers generally to a substance that is capable of converting the contact product of: 1) a metallocene component; and 2) a component that provides an activatable ligand such as an alkyl or hydride ligand to the metallocene, when the metallocene compound does not already comprise such a ligand; into a catalyst that can polymerize olefins. This term is used regardless of whether an activator ionizes the metallocene, abstracts an anionic ligand to form an ion pair, weakens a metal-ligand bond in the metallocene, simply coordinates to an anionic ligand, or any other mechanism. As disclosed herein, the contact product comprises at least one activator, which can be selected independently from: i) an activator-support comprising a solid oxide treated with an electron-withdrawing anion, a layered mineral, an ion-exchangeable activator-support, or any combination thereof; ii) an organoaluminoxane compound; iii) an organoboron or an organoborate compound; or iv) any combination of these components.

[0152] The term "clay" is used herein to refer to that component of the catalyst composition, a substantial portion of which constitutes a clay mineral or a mixture of clay minerals that have been pretreated by either exchanging cations, pillaring or simply wetting, that can be used as a activator-support in the catalyst composition described herein. The transition metal compound and organometal cocatalyst are reacted with the clay activator-support to form the active catalyst. While not intending to be bound by the following statement, the clay component of the catalyst composition of this invention probably functions as a activator-support for the transition metal compound, as well as a cocatalyst from the standpoint that it is in intimate physical chemical contact with the transition metal component.

[0153] As used herein, the collective term "clay mineral" is used herein to describe the large group of finely-crystalline, sheet like clay minerals that are found in nature in fine-grained sediments, sedimentary rocks, and the like. Clay minerals are a class of hydrous silicate and aluminosilicate minerals with sheet-like structures and very high surface areas. This term is also used to describe hydrous magnesium silicates with a phyllosilicate structure. Many common clay minerals belong to the kaolinite, montmorillonite, or illite groups of clays. Thus, the term "clay mineral" is not used herein to refer to the fine-grained soil consisting of mineral particles, not necessarily clay minerals, that are less than 0.002 mm in size.

[0154] The term "pillared clay" is used herein to refer to a component of the catalyst composition comprising clay minerals, typically of the of the smectite group and other phyllosilicates in addition to sepiolites and palygorskites, that have been ion exchanged with large, typically polynuclear, highly charged metal complex cations. Examples of such ions include Keggin ions which can have charges such as 7+, various polyoxometallates, and other large ions. Thus, the term pillaring refers to a simple exchange reaction in which the exchangeable cations of a clay material are replaced with large, highly charged ions, such as Keggin ions. These polymeric cations are then immobilized within the interlayers of the clay and when calcined are converted to metal oxide "pillars," effectively supporting the clay layers as column-like structures. Examples of pillaring and pillared clays are found in: T.J. Pinnavaia, *Science* 220 (4595), 365-371 (1983); J.M. Thomas, *Intercalation Chemistry*, (S. Whittington and A. Jacobson, eds.) Ch. 3, pp. 55-99, Academic Press, Inc., (1972); U.S. Patent No. 4,452,910; U.S. Patent No. 5,376,611; and U.S. Patent No. 4,060,480.

[0155] The publications discussed above and throughout the text are provided solely for their disclosure prior to the filing date of the present application. Nothing herein is to be construed as an admission that the inventors are not entitled to antedate such disclosure by virtue of prior invention.

[0156] For any particular compound disclosed herein, any general structure presented also encompasses all conformational isomers, regioisomers, stereoisomers, and the like, that can arise from a particular set of substituents. The general structure also encompasses all enantiomers, diastereomers, and other optical isomers whether in enantiomeric or racemic forms, as well as mixtures of stereoisomers, as the context requires.

[0157] The present invention is further illustrated by the following examples, which are not to be construed in any way as imposing limitations upon the scope thereof.

[0158] In the following examples, unless otherwise specified, the syntheses and preparations described herein were carried out under an inert atmosphere such as nitrogen and/or argon. Solvents were purchased from commercial sources

and were typically dried over activated alumina prior to use. Unless otherwise specified, reagents were obtained from commercial sources.

[0159] General testing procedures, characterization, and synthetic procedures are provided herein. Synthetic methods to prepare the metallocenes, treated solid oxides, and other reagents of this invention are also provided herein.

General Testing Procedures

[0160] Melt index (MI, g/10 min) was determined in accordance with ASTM D1238 condition F at 190°C with a 2,160 gram weight.

[0161] High load melt index (HLMI, g/10 min) was determined in accordance with ASTM D1238 condition E at 190°C with a 21,600 gram weight.

[0162] Polymer density was determined in grams per cubic centimeter (g/cc) on a compression molded sample, cooled at about 15°C per hour, and conditioned for about 40 hours at room temperature in accordance with ASTM D1505 and ASTM D1928, procedure C.

[0163] Molecular weight and molecular weight distributions were obtained using a PL-GPC 220 (Polymer Labs, UK) system equipped with a differential refractive index detector and three 7.5 mm x 300 mm 20 µm Mixed A-LS columns (Polymer Labs) running at 145 °C. The flow rate of the mobile phase, 1,2,4-trichlorobenzene (TCB) containing 0.5 g/L 2,6-di-*t*-butyl-4-methylphenol (BHT), was set at 1 mL/min and the concentration of polymer solutions was generally kept in the range of 1.0-1.5 mg/mL, depending on the molecular weights. Sample preparation was conducted at 150 °C for 4 h with occasional and gentle agitation before the solutions being transferred to sample vials for injection. In order to minimize unbalanced solvent peak, solvent with the same composition as the mobile phase was used for solution preparation. The integral calibration method was employed to deduce molecular weights and molecular weight distributions using a Chevron Phillips Chemicals Company's broad linear polyethylene, Marlex BHB5003, as the broad standard. The integral table of the broad standard was pre-determined in a separate experiment with SEC-MALS.

Melt viscosity measurements. to determine shear viscosity characteristics

[0164] Small-strain oscillatory shear measurements were performed on an ARES oscillatory rheometer using parallel-plate geometry (TA Instruments, formerly Rheometrics Inc.). Data were typically obtained over an angular frequency range of 0.03 to 100 rad/s at a temperature of 190 °C.

[0165] Fluff samples were stabilized with 0.1 wt% BHT dispersed in acetone and then vacuum dried before molding. Samples were compression molded at 184 °C for a total of three minutes. The samples were allowed to melt at a relatively low pressure for one minute and then subjected to a high molding pressure for an additional two minutes. The molded samples were then quenched in a cold (room temperature) press. Disks having the size 2 mm x 25.4 mm diameter were stamped out of the molded slabs for rheological characterization.

[0166] The test chamber of the rheometer was blanketed in nitrogen in order to minimize polymer degradation. The rheometer was preheated to the initial temperature of the study. Upon sample loading and after oven thermal equilibration, the specimens were squeezed between the plates to a 1.6 mm thickness and the excess was trimmed.

[0167] Strains were generally maintained at a single value throughout a frequency sweep but larger strain values were used for low viscosity samples to maintain a measurable torque. Smaller strain values were used for high viscosity samples to avoid overloading the torque transducer and to keep within the linear viscoelastic limits of the sample. The instrument automatically reduces the strain at high frequencies if necessary to keep from overloading the torque transducer.

[0168] Viscosity data were fit with a modified Carreau-Yasuda model [R. Byron Bird, Robert C. Armstrong, and Ole Hassager, Dynamics of Polymeric Liquids, Volume 1, Fluid Mechanics, (John Wiley & Sons, New York, 1987), p 171-172], to obtain estimates of the zero shear viscosity, viscous relaxation time, and a breadth parameter, as indicated below.

$$|\eta^*| = \eta_0 / [1 + (\omega\tau_\eta)^a]^{(1-n)/a},$$

wherein:

$|\eta^*|$ = magnitude of the complex viscosity in Pa·s

ω = angular frequency in rad/s

η_0 = zero shear viscosity in Pa·s

τ_η = viscous relaxation time in s

a = breadth parameter

n = power law parameter, fixed at 0.1818.

Absolute Molecular Weight as Determined by Light Scattering

[0169] Molecular weight data were determined using SEC-MALS, which combines the methods of size exclusion chromatography (SEC) with multi-angle light scattering (MALS) detection. A DAWN EOS 18-angle light scattering photometer (Wyatt Technology, Santa Barbara, CA) was attached to a PL-210 SEC system (Polymer Labs, UK) or a Waters 150 CV *Plus* system (Milford, MA) through a hot transfer line, thermally controlled at the same temperature as the SEC columns and its differential refractive index (DRI) detector (145 °C). At a flow rate setting of 0.7 mL/min, the mobile phase, 1,2,4-trichlorobenzene (TCB), was eluted through three, 7.5 mm x 300 mm, 20 µm Mixed A-LS columns (Polymer Labs). Polyethylene (PE) solutions with concentrations of ~1.2 mg/mL, depending on samples, were prepared at 150 °C for 4 h before being transferred to the SEC injection vials sitting in a carousel heated at 145 °C. For polymers of higher molecular weight, longer heating times were necessary in order to obtain true homogeneous solutions. In addition to acquiring a concentration chromatogram, seventeen light-scattering chromatograms at different angles were also acquired for each injection using Wyatt's Astra® software. At each chromatographic slice, both the absolute molecular weight (M) and root mean square (RMS) radius, also known as radius of gyration (R_g) were obtained from a Debye plot's intercept and slope, respectively. Methods for this process are detailed in Wyatt, P.J., *Anal. Chim. Acta*, 272, 1 (1993). The linear PE control employed was a linear, high-density broad molecular weight distribution (MWD) polyethylene sample (Chevron Phillips Chemical Co.). The weight average molecular weight (M_w), number average molecular weight (M_n), z-average molecular weight (M_z) and molecular weight distribution (M_w/M_n) were computed from these data, and are presented in various Tables.

[0170] The Zimm-Stockmayer approach was used to determine the amount of LCB in ethylene polymers. Since SEC-MALS measures M and R_g at each slice of a chromatogram simultaneously, the branching indices, g_M , as a function of M could be determined at each slice directly by determining the ratio of the mean square R_g of branched molecules to that of linear ones, at the same M , as shown in equation 1:

$$g_M = \frac{\langle R_g \rangle_{br}^2}{\langle R_g \rangle_{lin}^2} \quad (1)$$

where the subscripts *br* and *lin* represent branched and linear polymers, respectively.

[0171] At a given g_M , the weight-averaged number of LCB per molecule (B_{3w}) was computed using Zimm-Stockmayer's equation, shown in equation 2, where the branches were assumed to be trifunctional, or Y-shaped.

$$g_M = \frac{6}{B_{3w}} \left\{ \frac{1}{2} \left(\frac{2 + B_{3w}}{B_{3w}} \right)^{1/2} \ln \left[\frac{(2 + B_{3w})^{1/2} + (B_{3w})^{1/2}}{(2 + B_{3w})^{1/2} - (B_{3w})^{1/2}} \right] - 1 \right\} \quad (2)$$

LCB frequency (LCB_{Mi}), the number of LCB per 1 000 C, of the i^{th} slice was then computed straightforwardly using equation 3:

$$LCB_{Mi} = 1\,000 * 14 * B_{3w} / M_i \quad (3)$$

where M_i is the MW of the i^{th} slice. The LCB distribution (LCBD) across the molecular weight distribution (MWD) was thus established for a full polymer.

[0172] A Quantachrome Autosorb-6 Nitrogen Pore Size Distribution Instrument was used to determine specific surface area ("surface area") and specific pore volume ("pore volume"). This instrument was acquired from the Quantachrome Corporation, Syosset, N.Y.

Preparation of a Fluorided Silica-Alumina Activator-Support

[0173] The silica-alumina used to prepare the fluorided silica-alumina acidic activator-support in this Example was

typically Davison silica-alumina obtained from W.R. Grace as Grade MS13-110, containing 13% alumina, having a pore volume of about 1.2 cc/g and a surface area of about 400 m²/g. This material was fluorided by impregnation to incipient wetness with a solution containing ammonium bifluoride in an amount sufficient to equal 10 wt % of the weight of the silica-alumina. This impregnated material was then dried in a vacuum oven for 8 hours at 100°C. The thus-fluorided silica-alumina samples were then calcined as follows. About 10 grams of the alumina were placed in a 1.75-inch quartz tube fitted with a sintered quartz disk at the bottom. While the silica was supported on the disk, dry air was blown up through the disk at the linear rate of about 1.6 to 1.8 standard cubic feet per hour. An electric furnace around the quartz tube was used to increase the temperature of the tube at the rate of about 400°C per hour to a final temperature of about 500°C. At this temperature, the silica-alumina was allowed to fluidize for about three hours in the dry air. Afterward, the silica-alumina was collected and stored under dry nitrogen, and was used without exposure to the atmosphere.

Preparation of a Sulfated Alumina Activator-Support

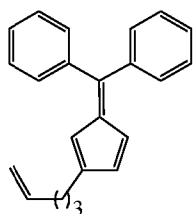
[0174] Sulfated alumina was formed by a process wherein alumina was chemically-treated with a sulfate or bisulfate source, typically selected from, but not limited to, sulfuric acid, ammonium sulfate, or ammonium bisulfate. One example follows.

[0175] A commercial alumina sold as W.R. Grace Alumina A was sulfated by impregnation with an aqueous solution containing about 15-20% (NH₄)₂SO₄ or H₂SO₄. This sulfated alumina was calcined at 550 °C in air (240 °C/h ramp rate), with a 3 h hold period at this temperature. Afterward, the alumina was collected and stored under dry nitrogen, and was used without exposure to the atmosphere.

Metallocene Preparations

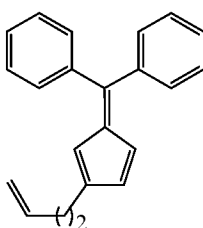
[0176] Unless specified otherwise, reagents were obtained from Aldrich Chemical Company and were used as received. 2,7-Di-tert-butylfluorene was purchased from Degussa. The Grignard reagent CpMgCl (1M in THF) was purchased from Boulder Scientific Company. Hafnium(IV) chloride was purchased from Strem. The solvent THF was distilled from potassium, while anhydrous diethyl ether, methylene chloride, pentane, and toluene were purchased from Fisher Scientific Company and stored over activated alumina. All solvents were degassed and stored under nitrogen. Reaction products were analyzed by ¹H NMR spectroscopy (300 MHz, CDCl₃ referenced against either the residual proton peak at 7.24 ppm for CHCl₃ or TMS at 0 ppm) or ¹³C NMR (75 MHz, CDCl₃, referenced against central line of CDCl₃ at 77.00 ppm). Reported preparations were not optimized.

[0177] The following fulvenes, **F-1** through **F-5**, were prepared as disclosed herein and used to prepare the ligands **L-1** through **L-5** provided herein.



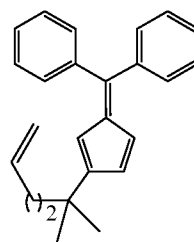
2-(pent-4-enyl)-6,6-diphenylpentafulvene

F-1



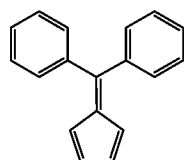
2-(but-3-enyl)-6,6-diphenylpentafulvene

F-2



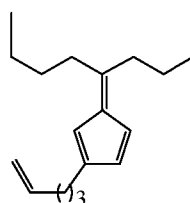
2-(1,1-dimethylpent-4-enyl)-6,6-diphenylpentafulvene

F-3



6,6-diphenylpentafulvene

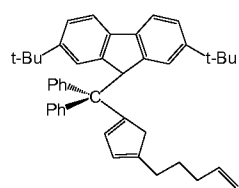
F-4



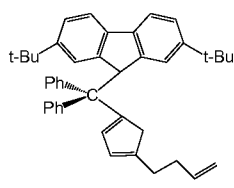
2-(pent-4-enyl)-6,6-dibutylpentafulvene

F-5

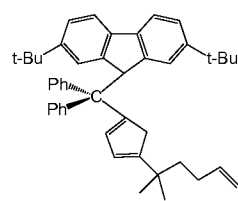
[0178] The following ligands L-1 through L-5 were prepared as disclosed herein.



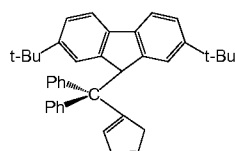
Mixture of isomers

L-1

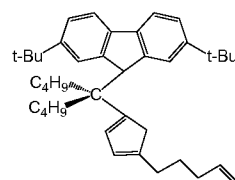
Mixture of isomers

L-2

Mixture of isomers

L-3

Mixture of isomers

L-4

Mixture of isomers

L-5

Synthesis of 2-(pent-4-enyl)-6,6-diphenylpentafulvene (F-1).

[0179] To 5-bromo-1-pentene (100 g of 95 wt%, 0.637 mol) was added cyclopentadienyl magnesium chloride (700 mL of 1 M solution in THF, 0.7 mol) at 0 °C in an hour. After stirring for an additional 30 minutes at 0 °C, the mixture was warmed to room temperature. After stirring overnight, the reaction was quenched with a mixture of ice and water. The mixture was extracted with pentane. The organic layer was washed with water and dried over anhydrous sodium sulfate. Removal of the solvent under vacuum at room temperature gave a yellow-brown liquid (98 g, crude pent-4-enylcyclopentadiene). To the crude pent-4-enylcyclopentadiene (89 g) dissolved in THF (500 mL) was added n-BuLi (60 mL of 10 M in hexanes, 0.6 mol) at -78 °C. The mixture was warmed up to room temperature and stirred overnight. The anion solution was added to benzophenone (110 g, 0.604 mol) dissolved in THF (500 mL) at 0 °C in 25 minutes. The mixture was warmed to room temperature and stirred overnight. The reaction was quenched with a mixture of ice and 10% HCl aqueous solution. The mixture was extracted with pentane. The organic layer was washed with water and dried over anhydrous sodium sulfate. Removal of the solvent under vacuum at 40 °C gave a dark red viscous oil. The oil was dissolved in heptane and filtered through silica gel. The product was collected by washing the silica gel with 5-10% CH₂Cl₂ in heptane. Removal of the solvent gave the desired product (145 g, 84% yield based on 5-bromo-1-pentene) as a dark red viscous oil. ¹H NMR (300 MHz, CDCl₃) δ 7.41-7.48 (m, 10H), 6.59-6.62 (dd, J = 5.1 Hz, 1.4 Hz, 1H), 6.40-6.42 (dd, J = 5.1 Hz, 1.4 Hz, 1H), 6.12-6.15 (m, 1H), 5.86-6.02 (m, 1H), 5.08-5.20 (m, 2H), 2.55-2.60 (t, J = 7.2 Hz, 2H), 2.22-2.30 (m, 2H), 1.76-1.88 (quin, J = 7.2 Hz, 2H); ¹³C NMR (75 MHz, CDCl₃) δ 148.28, 148.13, 143.28, 140.85, 140.76, 138.01, 133.51, 131.34, 131.29, 127.76, 127.74, 127.13, 127.08, 124.74, 118.24, 114.24, 33.95, 30.13, 28.46.

Synthesis of 1-(3-(pent-4-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (L-1)

[0180] To 2,7-di-tert-butylfluorene (125.1 g, 0.45 mol) dissolved in Et₂O (700 mL) was added n-BuLi (47 mL of 10 M in hexanes, 0.47 mol) at 0 °C. The mixture was warmed to room temperature and stirred overnight. The anion solution was added to 2-(pent-4-enyl)-6,6-diphenylpentafulvene (F-1) (145 g, 0.487 mol) dissolved in Et₂O (300 mL) at -78 °C in 10 minutes. The mixture was warmed to room temperature and stirred overnight. The reaction was quenched with a mixture of ice and 10% HCl aqueous solution. The mixture was extracted with Et₂O. The organic layer was washed with water and dried over anhydrous sodium sulfate. Removal of the solvent under vacuum gave a pale brown solid. The solid was washed with heptane and dried under vacuum. A mixture of isomers for the desired product (191.7 g, 74% yield) was obtained as a white solid.

Synthesis of 2-(but-3-enyl)-6,6-diphenylpentafulvene (F-2)

[0181] To 4-bromo-1-butene (100 g of 97 wt%, 0.719 mol) was added cyclopentadienyl magnesium chloride (800 mL of 1 M solution in THF, 0.8 mol) at 0 °C in 50 minutes. After stirring for an additional 15 minutes at 0 °C, the mixture was warmed to room temperature. After stirring overnight, the reaction was quenched with a mixture of ice and water. The

mixture was extracted with pentane. The organic layer was washed with water and dried over anhydrous sodium sulfate. Removal of the solvent under vacuum at room temperature gave a brown liquid (94.2 g, crude but-3-enylcyclopentadiene). To the crude but-3-enylcyclopentadiene (94.2 g) dissolved in THF (500 mL) was added n-BuLi (70 mL of 10 M in hexanes, 0.7 mol at -78 °C. The mixture was warmed up to room temperature and stirred overnight. The anion solution was added to benzophenone (133.8 g, 0.735 mol) dissolved in THF (400 mL) at 0 °C in 35 minutes. The mixture was warmed to room temperature and stirred overnight. The reaction was quenched with a mixture of ice and 10% HCl aqueous solution. The mixture was extracted with pentane. The organic layer was washed with water and dried over anhydrous sodium sulfate. Removal of the solvent under vacuum at 40 °C gave a dark red viscous oil. The oil was dissolved in heptane and filtered through silica gel. The product was collected by washing the silica gel with 5-10% CH₂Cl₂ in heptane. Removal of the solvent gave the desired product (152g, 74.4% yield based on 4-bromo-1-butene) as a dark red viscous oil. ¹H NMR (300 MHz, CDCl₃) δ 7.29-7.41 (m, 10H), 6.50-6.53 (dd, *J* = 5.2 Hz, 1.4 Hz, 1H), 6.29-6.31 (dd, *J* = 5.2 Hz, 1.4 Hz, 1H), 6.02-6.05 (m, 1H), 5.82-5.98 (m, 1H), 4.94-5.16 (m, 2H), 2.53-2.60 (m, 2H), 2.33-2.43 (m, 2H); ¹³C NMR (75 MHz, CDCl₃) δ 148.59, 147.67, 143.18, 140.86, 140.78, 137.85, 133.48, 131.38, 131.36, 127.85, 127.82, 127.18, 127.13, 124.75, 118.35, 114.29, 33.36, 30.20.

Synthesis of 1-(3-(but-3-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (L-2)

[0182] To 2,7-di-tert-butylfluorene (91.7 g, 0.33 mol) dissolved in Et₂O (500 mL) was added n-BuLi (35 mL of 10 M in hexanes, 0.35 mol) at 0 °C. The mixture was warmed to room temperature and stirred overnight. The anion solution was added to 2-(but-3-enyl)-6,6-diphenylpentafulvene (compound F-2) (104 g, 0.366 mol) dissolved in Et₂O (200 mL) at 0 °C in 35 minutes. After stirring for an additional 30 minutes at 0 °C, the mixture was warmed to room temperature and stirred overnight. The reaction was quenched with a mixture of ice and 10% HCl aqueous solution. The mixture was extracted with CH₂Cl₂. The organic layer was washed with water and dried over anhydrous sodium sulfate. Removal of the solvent under vacuum gave a pale brown solid. The solid was washed with heptane and dried under vacuum. A mixture of isomers for the desired product (142 g, 76.5% yield) was obtained as a white solid.

Synthesis of 2-(1,1-dimethylpent-4-enyl)-6,6-diphenylpentafulvene (F-3)

[0183] To a solution of 6-butenyl-6-methylpentafulvene (17.8 g, 122 mmol) (prepared by the method of K. J. Stone and R. D. Little, J. Org. Chem., 1984, 49(11), 1849-1853) in dry THF (50 mL) was added a solution of methyllithium (75 mL of 1.6 M in ether, 120 mmol) while cooling in dry ice. After stirring for 20 hours and warming to room temperature, the yellow solution was gradually added to a solution of benzophenone (21.87 g, 120 mmol) in THF (50 mL) while cooling in ice. A red color formed immediately and after 4 hours analysis of an aliquot showed that the reaction was nearly complete. After an additional hour, the mixture was cooled while a solution of concentrated hydrochloric acid (20 mL) in water (200 mL) was added. Following addition of pentane (150 mL), the organic layer was washed with water and dried over sodium sulfate. The solvent was removed under vacuum and the red liquid was cooled to -15 °C overnight. The red crystalline product was washed with cold methanol and dried under vacuum to a red solid (32.8 g, 84% yield). ¹H NMR (300 MHz, CDCl₃) δ 7.22-7.40 (m, 10H), 6.56-6.58 (dd, *J* = 5.1 Hz, 1.8 Hz, 1H), 6.24-6.26 (dd, *J* = 5.1 Hz, 1.8 Hz, 1H), 5.91-5.93 (t, *J* = 1.8 Hz, 1H), 5.70-5.85 (m, 1H), 4.84-5.00 (m, 2H), 1.88-2.00 (m, 2H), 1.52-1.60 (m, 2H), 1.17 (s, 6H); ¹³C NMR (75 MHz, CDCl₃) δ 156.16, 148.39, 143.20, 140.96, 140.92, 138.98, 131.61, 131.43, 131.39, 127.81, 127.77, 127.24, 127.14, 124.88, 116.30, 113.45, 41.96, 35.86, 29.90, 27.90.

Synthesis of 1-(3-(1,1-dimethylpent-4-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (L-3)

[0184] A solution of 2,7-di-tert-butylfluorene (27.8 g, 100 mmol) in Et₂O (200 mL) was cooled in dry ice and n-BuLi (68 mL of 1.6 M in hexanes, 0.109 mmol) was added dropwise. The slurry was warmed to room temperature and stirred for 24 hours. The dark solution was cooled in dry ice and a solution of 2-(1,1-dimethylpent-4-enyl)-6,6-diphenylpentafulvene (compound F-3) (32.8 g, 54.3 mmol) in Et₂O (100 mL) was then added rapidly. The mixture was warmed to room temperature and stirred for 20 hours. After cooling in ice, a solution of concentrated hydrochloric acid (20 mL) in water (200 mL) was added. Following addition of pentane (100 mL), the organic layer was separated and washed with water. After drying over sodium sulfate and filtering, the solvent was removed under vacuum leaving a glassy solid. The solid was heated with methanol (100 mL) and the hot methanol solution was poured off. This process was repeated four times. The solid was then dissolved in hot pentane, which was then removed under vacuum while heating. The solid was broken up, dried under vacuum, and then heated with ethanol (70 mL). After cooling, the solid was filtered and dried. A mixture of isomers for the desired product (18.1 g, 30% yield) was obtained as a white solid.

Synthesis of 6,6-diphenylpentafulvene (F-4)

[0185] Benzophenone (63.8 g, 350 mmol) was dissolved in anhydrous 1,2-dimethoxyethane (DME) (150 mL) under nitrogen. In a one-liter flask, ground potassium hydroxide (30 g, 535 mmol) was slurried in DME (200 mL). The slurry was cooled in an ice bath and freshly cracked cyclopentadiene (35 mL, 430 mmol) were added. After 30 minutes, the solution of benzophenone was added over 15 minutes. The flask was stirred in a refrigerator for 90 hours and then, while cooling in ice, 3M HCl (450 mL) were added. The mixture was diluted with pentane (500 mL) and separated. The organic layer was washed with water (2 X 200 mL) and dried over sodium sulfate. The solution was filtered and taken to dryness under vacuum. The solid was dissolved in boiling pentane (600 mL) and then concentrated to 400 mL. Cooling to -15 °C for 40 hours gave a red solid (69.5 g, 86.3% yield). ¹H NMR (300 MHz, CDCl₃) δ 7.24-7.38 (m, 10H), 6.53-6.59 (m, 2H), 6.24-6.30 (m, 2H); ¹³C NMR (75 MHz, CDCl₃) δ 151.24, 143.20, 140.65, 131.73, 131.55, 128.16, 127.20, 123.89.

Synthesis of 1-cyclopentadienyl-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (L-4)

[0186] To a solution of 2,7-di-tert-butylfluorene (29.8 g, 107 mmol) in dry THF (100 mL), cooled in dry ice, was added n-BuLi (43.0 mL of 2.5 M in hexanes, 107.5 mmol). The bath was removed and the dark solution was stirred for 2 hours. This solution was then added dropwise to a solution of 6,6-diphenylpentafulvene (compound F-4) (26.0 g, 113 mmol) in THF (100 mL), while cooling in ice. The reaction mixture was stirred at room temperature for 86 hours and then cooled in ice. 1M HCl solution, (100 mL) was added. The mixture was diluted with chloroform (100 mL) and separated. The chloroform layer was washed with water (3 X 100 mL) and dried over sodium sulfate. The solution was filtered and evaporated to a light orange solid. The solid was dissolved in boiling chloroform (150 mL) and methanol (150 mL) was slowly added. After cooling for two days to -15 °C, the solid was filtered off, ground, and dried under vacuum. A mixture of isomers for the desired product (25.4 g, 46.7% yield) was obtained as an off white solid.

Synthesis of 5-(3-(pent-4-enyl)cyclopentadienyl)-5-(2,7-di-tert-butylfluorenyl)nonane (L-5)

[0187] A flask was charged with 2,7-di-tert-butylfluorene (10 g, 36 mmol), Et₂O (150 mL), cooled to -78 °C under N₂ and stirred as n-BuLi (4.3 mL of 10 M in hexanes, 43 mmol) was added via syringe. The reaction mixture was warmed to room temperature, stirred overnight, cooled to -78 °C, and 2-(pent-4-enyl)-6,6-dibutylpentafulvene (compound F-5) (13 g, 50 mmol) (prepared by the method of K.J. Stone and R.D. Little, J. Org. Chem., (1984), 49(11), 1849-1853) quickly added. The reaction mixture was warmed to room temperature and stirred overnight. The reaction was quenched with saturated NH₄Cl solution. The organic layer was extracted with Et₂O, washed with water and dried over anhydrous Na₂SO₄. Removal of the solvent under vacuum gave a yellow oil. Elution of the oil through a silica-gel column using heptane afforded a mixture of isomers for the desired product (12.8 g, 66% yield) as an oil.

Synthesis of diphenylmethylidene{η⁵-[3-(pent-4-enyl)cyclopentadien-1-ylidene]}[η⁵-(2,7-di-tert-butylfluoren-9-ylidene)]hafnium dichloride (I-1)

[0188] To 1-(3-(pent-4-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (compound L-1) (45.3 g, 78.6 mmol) dissolved in Et₂O (400 mL) was slowly added n-BuLi (68.5 mL of 2.5 M in hexanes, 171.3 mmol) at 0 °C. The mixture was warmed to room temperature, stirred overnight, and then added via cannula to HfCl₄ (26.8 g, 83.6 mmol) suspended in a mixture of pentane (450 mL) and Et₂O (30 mL) at 0 °C in 30 minutes. The mixture was warmed to room temperature and stirred for two days. The slurry was concentrated and centrifuged. The liquid was decanted off. The remaining solid was washed a second time with pentane (100 mL), then extracted with methylene chloride and centrifuged. The solution was taken to dryness under vacuum to give a yellow solid (46.4 g, 71.7%). ¹H NMR (300 MHz, CDCl₃) δ 7.88-7.98 (m, 3H), 7.78-7.88 (m, 3H), 7.40-7.50 (m, 2H), 7.29-7.38 (broad t, J = 7.2 Hz, 2H), 7.11-7.28 (m, 4H), 6.28 (broad s, 1H), 6.24 (broad s, 1H), 5.87-5.93 (t, J = 2.7 Hz, 1H), 5.61-5.78 (m, 1H), 5.44-5.50 (t, J = 2.7 Hz, 1H), 5.19-5.25 (t, J = 2.7 Hz, 1H), 4.82-4.96 (m, 2H), 2.28-2.48 (m, 2H), 1.94-2.05 (m, 2H), 1.46-1.60 (m, 2H), 0.98 (s, 18H); ¹³C NMR (75 MHz, CDCl₃) δ 149.41, 149.21, 144.47, 144.24, 137.71, 132.69, 129.08, 128.83, 128.45, 128.39, 128.22, 126.50, 126.46, 126.13, 125.97, 123.70, 123.46, 123.40, 123.34, 119.89, 119.66, 119.01, 118.86, 118.82, 118.53, 114.75, 114.39, 111.11, 100.92, 100.69, 76.88, 57.88, 35.29, 35.27, 33.75, 31.04, 31.02, 29.48, 29.31.

Synthesis of diphenylmethylidene{η⁵-[3-(but-3-enyl)cyclopentadien-1-ylidene]}[η⁵-(2,7-di-tert-butylfluoren-9-ylidene)] hafnium dichloride (1-2)

[0189] To 1-(3-(but-3-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (compound L-2) (3.2 g, 5.7 mmol) dissolved in Et₂O (30 mL) was slowly added n-BuLi (5.2 mL of 2.5 M in hexanes, 13 mmol) at 0 °C. The mixture was warmed to room temperature, stirred overnight, and then added via cannula to HfCl₄ (2.1 g, 6.5 mmol)

suspended in a mixture of pentane (30 mL) and Et₂O (5 mL) at 0 °C in 10 minutes. The mixture was warmed to room temperature and stirred for two days. The slurry was concentrated and centrifuged. The liquid was decanted off. The remaining solid was washed a second time with pentane (80 mL), then extracted with methylene chloride and centrifuged. The solution was taken to dryness under vacuum to give a yellow solid (3.1 g, 67.4% yield). ¹H NMR (300 MHz, CDCl₃) δ 7.87-7.98 (m, 3H), 7.79-7.86 (m, 3H), 7.43-7.49 (m, 2H), 7.30-7.38 (dt, *J* = 7.5 Hz, 1.4 Hz, 2H), 7.14-7.29 (m, 4H), 6.24-6.27 (d, *J* = 0.6 Hz, 1H), 6.20-6.24 (d, *J* = 0.6 Hz, 1H), 5.87-5.92 (t, *J* = 2.7 Hz, 1H), 5.62-5.77 (m, 1H), 5.42-5.47 (t, *J* = 2.7 Hz, 1H), 5.18-5.23 (t, *J* = 2.7 Hz, 1H), 4.85-4.98 (m, 2H), 2.35-2.55 (m, 2H), 2.13-2.22 (m, 2H), 0.96 (s, 18H); ¹³C NMR (75 MHz, CDCl₃) δ 149.52, 149.33, 144.51, 144.30, 137.33, 132.16, 129.13, 128.89, 128.51, 128.45, 128.30, 128.26, 126.58, 126.53, 126.24, 126.06, 123.77, 123.54, 123.42, 123.36, 119.97, 119.75, 119.08, 118.90, 118.58, 114.94, 114.83, 111.14, 101.01, 100.68, 76.93, 57.94, 35.36, 35.35, 34.11, 31.08, 31.05, 29.42.

Synthesis of diphenylmethylidene{η⁵-[3-(but-3-enyl)cyclopentadien-1-ylidene]}[η⁵-(2,7-di-tert-butylfluoren-9-ylidene)]zirconium dichloride (I-3)

[0190] To 1-(3-(but-3-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (compound **L-2**) (40.5 g, 72.1 mmol) suspended in Et₂O (400 mL) was slowly added n-BuLi (15.2 mL of 10 M in hexanes, 152 mmol) at 0 °C. The mixture was warmed to room temperature, stirred overnight, and then added via cannula to ZrCl₄ (18.5 g, 79.4 mmol) suspended in a mixture of pentane (400 mL) and Et₂O (30 mL) at 0 °C in 15 minutes. The mixture was warmed to room temperature, stirred for one day, and evacuated to dryness. The residue was stirred in pentane (300 mL) and centrifuged. The supernatant was discarded. The remaining solid was washed a second time with pentane (100 mL), then extracted with methylene chloride and centrifuged. The solution was taken to dryness under vacuum to give a red solid (38.1 g, 73.3% yield). ¹H NMR (300 MHz, CDCl₃) δ 7.88-8.02 (m, 3H), 7.77-7.88 (m, 3H), 7.46-7.54 (m, 2H), 7.31-7.40 (broad t, *J* = 7.5 Hz, 2H), 7.14-7.32 (m, 4H), 6.24 (s, 1H), 6.20 (s, 1H), 5.96-6.02 (unresolved t, 1H), 5.63-5.79 (m, 1H), 5.50-5.55 (unresolved t, 1H), 5.25-5.31 (unresolved t, 1H), 4.87-5.01 (m, 2H), 2.33-2.53 (m, 2H), 2.15-2.27 (m, 2H), 0.97 (s, 18H); ¹³C NMR (75 MHz, CDCl₃) δ 149.85, 149.65, 144.23, 144.01, 137.27, 133.51, 129.08, 128.84, 128.50, 128.45, 128.33, 128.30, 126.58, 126.54, 126.18, 126.01, 124.04, 123.81, 123.55, 123.48, 121.08, 120.89, 120.31, 120.03, 119.43, 119.24, 115.71, 114.86, 108.44, 103.37, 103.18, 76.66, 58.10, 35.38, 35.36, 33.98, 31.05, 31.02, 29.46.

Synthesis of diphenylmethylidene{η⁵-[3-(pent-4-enyl)cyclopentadien-1-ylidene]}[η⁵-(2,7-di-tert-butylfluoren-9-ylidene)]zirconium dichloride (I-4)

[0191] To 1-(3-(pent-4-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (compound **L-1**) (34.7 g, 60.2 mmol) dissolved in Et₂O (300 mL) was slowly added n-BuLi (52 mL of 2.5 M in hexanes, 130 mmol) at 0 °C. The mixture was warmed to room temperature, stirred overnight, and then added via cannula to ZrCl₄ (14.7 g, 63.1 mmol) suspended in a mixture of pentane (250 mL) and Et₂O (20 mL) at 0 °C in 30 minutes. The mixture was warmed to room temperature, stirred for one day, and evacuated to dryness. The residue was stirred in pentane (200 mL) and centrifuged. The supernatant was discarded. The remaining solid was washed a second time with pentane (50 mL), then extracted with methylene chloride and centrifuged. The solution was taken to dryness under vacuum to give a red solid (33.5 g, 75.6%). ¹H NMR (300 MHz, CDCl₃) δ 7.94-7.99 (m, 2H), 7.89-7.94 (m, 1H), 7.77-7.87 (m, 3H), 7.47-7.53 (m, 2H), 7.32-7.39 (dt, *J* = 7.2 Hz, 1.2 Hz, 2H), 7.15-7.29 (m, 4H), 6.23 (broad s, 1H), 6.19 (broad s, 1H), 5.94-5.98 (t, *J* = 2.7 Hz, 1H), 5.62-5.76 (m, 1H), 5.50-5.54 (t, *J* = 2.7 Hz, 1H), 5.24-5.29 (t, *J* = 2.7 Hz, 1H), 4.82-4.96 (m, 2H), 2.23-2.43 (m, 2H), 1.97-2.05 (m, 2H), 1.48-1.61 (m, 2H), 0.97 (s, 18H); ¹³C NMR (75 MHz, CDCl₃) δ 149.85, 149.65, 144.27, 144.03, 137.79, 134.18, 129.11, 128.85, 128.51, 128.46, 128.34, 126.59, 126.55, 126.18, 126.03, 124.04, 123.79, 123.54, 123.47, 121.09, 120.89, 120.32, 120.06, 119.46, 119.26, 115.61, 114.44, 108.51, 103.36, 103.29, 76.69, 58.13, 35.39, 35.37, 33.78, 31.06, 31.03, 29.61, 29.33.

Synthesis of diphenylmethylidene{η⁵-[3-(1,1-dimethylpent-4-enyl)cyclopentadien-1-ylidene]}[η⁵-(2,7-di-tert-butylfluoren-9-ylidene)] zirconium dichloride (I-5)

[0192] A slurry of 1-(3-(1,1-dimethylpent-4-enyl)cyclopentadienyl)-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (compound **L-3**) (10.8 g, 17.9 mmol) in Et₂O (50 mL) was cooled in dry ice and n-BuLi (22.2 mL of 1.6 M in hexanes, 35.5 mmol) was added dropwise. After 1 hour, the bath was removed and the mixture was stirred for 48 hours at room temperature. The mixture was added to ZrCl₄ (4.37 g, 18.8 mmol) suspended in pentane (50 mL) while cooling in ice. The slurry was stirred for 65 hours at room temperature. The slurry was concentrated until thick and pentane (70 mL) was added. The slurry was stirred overnight and the liquid was decanted off. The solid was washed a second time with pentane and then extracted with methylene chloride and centrifuged. The solution was taken to dryness under vacuum to give a red solid (11.65 g, 85.2% yield). ¹H NMR (300 MHz, CDCl₃) δ 7.93-8.02 (m, 3H), 7.80-7.91 (m, 3H), 7.52-7.60 (dt, *J* = 8.7 Hz, 1.5 Hz, 2H), 7.38-7.47 (m, 2H), 7.20-7.35 (m, 4H), 6.27 (broad s, 2H), 6.14-6.18 (t, *J* = 3.0 Hz, 1H),

5.67-5.83 (m, 1H), 5.61-5.64 (t, $J = 3.0$ Hz, 1H), 5.48-5.52 (t, $J = 3.0$ Hz, 1H), 4.88-5.04 (m, 2H), 1.76-2.10 (m, 2H), 1.44-1.53 (m, 2H), 1.26 (s, 3H), 1.07 (s, 3H), 1.02 (s, 18H); ^{13}C NMR (75 MHz, CDCl_3) δ 149.67, 149.60, 144.31, 144.13, 143.46, 138.49, 129.15, 128.89, 128.51, 128.48, 128.39, 128.33, 126.58, 126.52, 126.11, 125.97, 124.18, 124.10, 123.73, 123.36, 121.09, 120.78, 120.20, 119.75, 118.88, 114.16, 113.84, 108.10, 104.30, 100.60, 77.19, 57.65, 46.43, 36.32, 35.38, 35.36, 31.06, 31.03, 29.47, 26.99, 24.19.

Reference Synthesis of Diphenylmethylidene[η^5 -(cyclopentadien-1-ylidene)][η^5 -(2,7-di-tert-butylfluoren-9-ylidene)]zirconium dichloride (C-1)

[0193] Under nitrogen, 1-cyclopentadienyl-1-(2,7-di-tert-butylfluorenyl)-1,1-diphenylmethane (compound **L-4**) (15.26 g, 30.0 mmol) was suspended in dry Et_2O (250 mL). While cooling in dry ice, $n\text{-BuLi}$ (24.0 mL of 2.5 M in hexanes, 60 mmol) were added dropwise. The bath was then removed and the mixture was stirred for 24 hours. The solution was gradually added to zirconium tetrachloride (7.38 g, 31.7 mmol) suspended in pentane (50 mL) and cooled in ice. The orange slurry was stirred for 90 hours and allowed to warm to room temperature. The resulting slurry was centrifuged and the solid was mixed with dry methylene chloride (120 mL). The mixture was centrifuged and the solution was removed and taken to dryness under vacuum. The desired product (9.63 g, 48% yield) was obtained as an orange solid. ^1H NMR (300 MHz, CDCl_3) δ 7.98-8.04 (d, $J = 9$ Hz, 2H), 7.91-7.96 (m, 2H), 7.83-7.89 (m, 2H), 7.55-7.60 (dd, $J = 9$ Hz, 1.8 Hz, 2H), 7.38-7.45 (dt, $J = 7.5$ Hz, 1.8 Hz, 2H), 7.21-7.36 (m, 4H), 6.30-6.34 (m, 4H), 5.64-5.67 (t, $J = 2.7$ Hz, 2H), 1.03 (s, 18H); ^{13}C NMR (75 MHz, CDCl_3) δ 149.98, 144.00, 128.93, 128.50, 128.41, 126.64, 126.08, 124.16, 123.56, 121.12, 120.30, 119.41, 117.92, 109.92, 102.40, 77.72, 58.36, 35.40, 31.01.

Reference Synthesis of dibutylmethylidene[η^5 -[3-(pent-4-enyl)cyclopentadien-1-ylidene]][η^5 -(2,7-di-tert-butylfluoren-9-ylidene)]zirconium dichloride (C-2)

[0194] A flask was charged with 5-(3-(pent-4-enyl)cyclopentadienyl)-5-(2,7-di-tert-butylfluorenyl)nonane (compound **L-5**) (12.8 g, 23.8 mmol), Et_2O (200 mL), stir bar and cooled to -78°C as $n\text{-BuLi}$ (5.3 mL of 10 M in hexanes, 53 mmol) was slowly added. The mixture was warmed to room temperature, stirred overnight, and then added via cannula to ZrCl_4 (5.5 g, 23.6 mmol) stirring in pentane at 0°C . The mixture was warmed to room temperature, stirred for 7 days, and evacuated to dryness. The residue was extracted with pentane, filtered, and the filtrate discarded. The remaining solid was extracted with CH_2Cl_2 , filtered and the filtrate evacuated to dryness affording a red solid (7.8 g, 47% yield). ^1H NMR (300 MHz, CDCl_3) δ 7.88-7.94 (m, 2H), 7.63 (broad s, 1H), 7.55 (broad s, 1H), 7.47-7.53 (m, 2H), 5.87-5.90 (t, $J = 2.7$ Hz, 1H), 5.58-5.73 (m, 1H), 5.46-5.49 (t, $J = 2.7$ Hz, 1H), 5.23-5.27 (t, $J = 2.7$ Hz, 1H), 4.80-4.92 (m, 2H), 2.55-2.75 (m, 4H), 2.20-2.40 (m, 2H), 1.90-2.00 (m, 2H), 1.40-1.80 (m, 10H), 1.15 (s, 18H), 1.00 (t, $J = 6.9$ Hz, 3H), 0.97 (t, $J = 6.9$ Hz, 3H); ^{13}C NMR (75 MHz, CDCl_3) δ 150.92, 150.64, 137.78, 134.97, 123.74, 123.51, 123.48, 123.44, 123.12, 122.48, 120.96, 120.56, 118.73, 118.36, 116.13, 114.32, 112.59, 102.07, 101.93, 76.73, 48.76, 35.79, 35.76, 34.81, 34.68, 33.75, 31.49, 31.48, 29.59, 29.24, 26.35, 26.18, 24.07, 24.04, 14.81, 14.78.

Reference Synthesis of dibutylmethylidene[η^5 -[3-(pent-4-enyl)cyclopentadien-1-ylidene]][η^5 -(2,7-di-tert-butylfluoren-9-ylidene)] hafnium dichloride (C-3)

[0195] To 5-(3-(pent-4-enyl)cyclopentadienyl)-5-(2,7-di-tert-butylfluorenyl)nonane (compound **L-5**) (14.6 g, 27.2 mmol) dissolved in Et_2O (150 mL) was slowly added $n\text{-BuLi}$ (26 mL of 2.5 M in hexanes, 65 mmol) at 0°C . The mixture was warmed to room temperature, stirred overnight, and then added via cannula to HfCl_4 (9.2 g, 28.7 mmol) suspended in a mixture of pentane (150 mL) and Et_2O (20 mL) at -78°C in 15 minutes. The mixture was warmed to room temperature, stirred for two days and evacuated to dryness. The residue was stirred in pentane (150 mL) and centrifuged. The supernatant was discarded. The remaining solid was extracted with methylene chloride and centrifuged. The solution was taken to dryness under vacuum to give a yellow solid (6.6 g, 31% yield). ^1H NMR (300 MHz, CDCl_3) δ 7.88-7.91 (m, 2H), 7.64 (broad s, 1H), 7.56 (broad s, 1H), 7.42-7.48 (m, 2H), 5.80-5.84 (t, $J = 2.7$ Hz, 1H), 5.58-5.73 (m, 1H), 5.39-5.43 (t, $J = 2.7$ Hz, 1H), 5.18-5.23 (t, $J = 2.7$ Hz, 1H), 4.78-4.91 (m, 2H), 2.55-2.75 (m, 4H), 2.22-2.42 (m, 2H), 1.88-1.97 (m, 2H), 1.40-1.80 (m, 10H), 1.27 (s, 18H), 0.99 (t, $J = 6.9$ Hz, 3H), 0.97 (t, $J = 6.9$ Hz, 3H); ^{13}C NMR (75 MHz, CDCl_3) δ 150.63, 150.32, 137.78, 133.55, 123.42, 123.40, 123.36, 123.15, 121.99, 121.30, 119.48, 119.02, 118.30, 117.97, 115.26, 115.07, 114.30, 99.44, 99.39, 76.79, 48.73, 35.74, 35.72, 35.01, 34.89, 33.78, 31.52, 31.49, 29.53, 29.30, 26.46, 26.30, 24.09, 24.05, 14.83, 14.79.

EXAMPLES 1-16

Catalytic Runs Varying the Metallocene, Activator-Support, and Conditions

5 **[0196]** Examples 1-16 in Table 1 illustrate ethylene polymerization runs performed in a one-gallon (3.785 liter) stainless steel autoclave reactor at various temperatures, using two liters of isobutane diluent and an aluminum alkyl cocatalyst and scavenger. No hydrogen or comonomer was added. Metallocene solutions (2mg/mL) were typically prepared by dissolving 30 mg of the metallocene in 15 mL of toluene. A typical polymerization procedure is as follows. The aluminum alkyl compound, treated solid oxide, and the metallocene solution were added through a charge port, typically in that
10 order, while venting isobutane vapor. The charge port was closed and two liters of isobutane were added. The contents of the reactor were stirred and heated to the desired run temperature (Table 1). Ethylene was fed on demand to maintain the specified pressure for the specified length of the polymerization run. The reactor was maintained at the desired run temperature through the run by an automated heating and cooling system.

15 **[0197]** After the allotted polymerization time, the ethylene flow was stopped, and the reactor slowly depressurized and opened to recover a granular polymer. In all cases, the reactor was clean with no indication of any wall scale, coating or other forms of fouling. The polymer was then removed and weighed (Table 1).

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Table 1. Polymerization data for polymerizations performed in autoclave reactor.

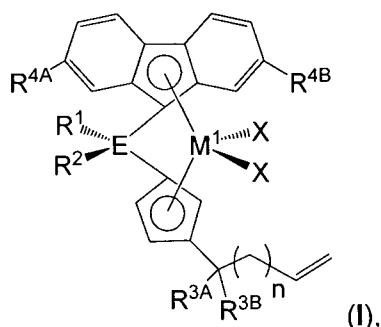
Example	Example Type	Metallocene	Time (min)	Temp (C)	Reactor pressure (kPa) (psi)	Support-Activator	Support-Activator weight (mg)	R3Al (mmol)	Metallocene weight (mg)	Solid PE (g)	Mn/1000	Mw/1000	Mz/1000	HI
1	Inventive	I-1	60	90	3102.6 (450)	Sulfated Alumina	100	0.5 TIBA	3.0	294	296	911	2187	3.07
2	Inventive	I-1	35	105	3102.6 (450)	Sulfated Alumina	100	0.2 TIBA	3.0	203	239	730	1877	3.06
3	Inventive	I-2	60	90	3102.6 (450)	Sulfated Alumina	100	0.5 TIBA	3.0	252	315	972	2258	3.09
4	Inventive	I-2	33	105	3102.6 (450)	Sulfated Alumina	100	0.2 TIBA	3.0	186	318	843	1942	2.65
5	Inventive	I-3	30	80	3102.6 (450)	Sulfated Alumina	100	0.25 TIBA	1.0	315	305	754	1492	2.47
6	Inventive	I-3	30	90	3102.6 (450)	Sulfated Alumina	105	0.5 TIBA	1.0	295	263	639	1223	2.43
7	Inventive	I-3	30	90	3102.6 (450)	Sulfated Alumina	104	0.5 TnBA	1.0	320	278	708	1383	2.55
8	Inventive	I-4	30	90	3102.6 (450)	Sulfated Alumina	101	0.5 TIBA	1.0	272	223	591	1221	2.65
9	Inventive	I-4	30	90	3102.6 (450)	Sulfated Alumina	108	0.5 TnBA	1.0	211	314	750	1429	2.39
10	Inventive	I-5	60	90	3102.6 (450)	Sulfated Alumina	50	0.5 TnBA	2.0	158	311	772	1575	2.48
11	Inventive	I-5	30	90	3102.6 (450)	Sulfated Alumina	100	0.25 TIBA	2.0	255	205	637	1270	3.11
12	Comparative	C-1	60	90	3102.6 (450)	Sulfated Alumina	100	0.25 TIBA	2.0	42	insolubles	insolubles	insolubles	N/A
13	Comparative	C-1	60	105	3102.6 (450)	Sulfated Alumina	100	0.25 TIBA	2.0	63	insolubles	insolubles	insolubles	N/A

(continued)

Example	Example Type	Metallocene	Time (min)	Temp (C)	Reactor pressure (kPa) (psi)	Support-Activator	Support-Activator weight (mg)	R3Al (mmol)	Metallocene weight (mg)	Solid PE (g)	Mn/1000	Mw/1000	Mz/1000	HI
14	Comparative	C-2	16	90	3102.6 (450)	Sulfated Alumina	50	0.5 TIBA	1.0	232	70	183	348	2.60
15	Comparative	C-3	60	90	3102.6 (450)	Sulfated Alumina	100	0.5 TIBA	3.0	294	108	375	1068	3.49
16	Comparative	C-3	60	100	3102.6 (450)	Sulfated Alumina	100	0.5 TIBA	3.0	369	97	267	640	2.75

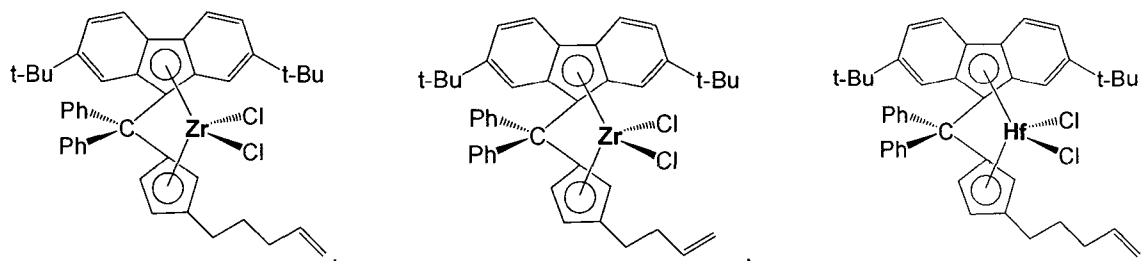
Claims

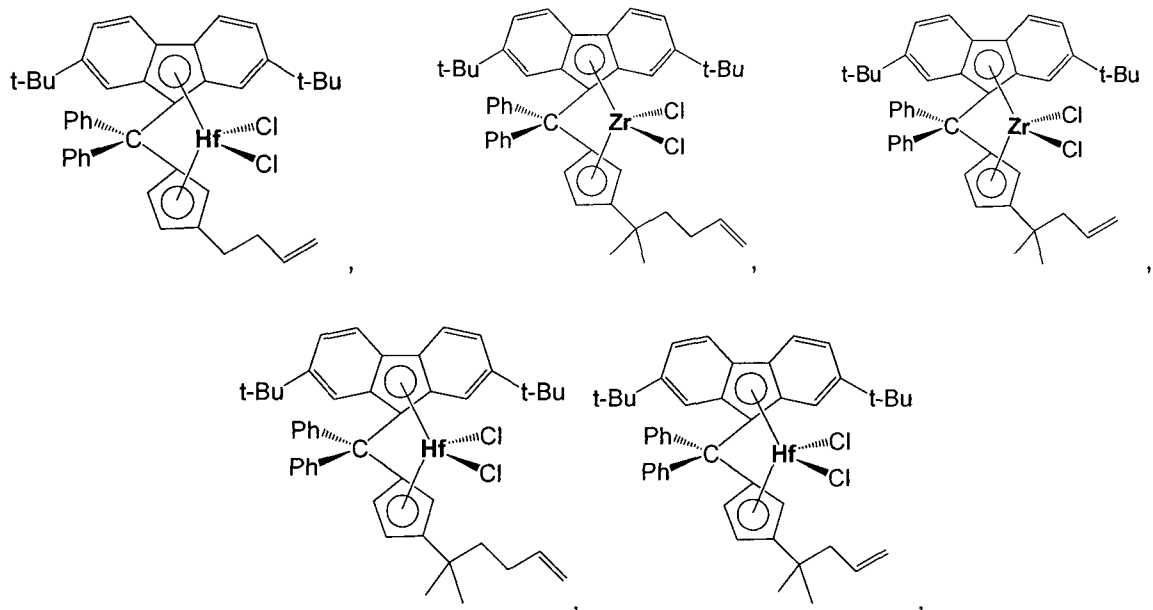
1. A compound having the formula:



wherein

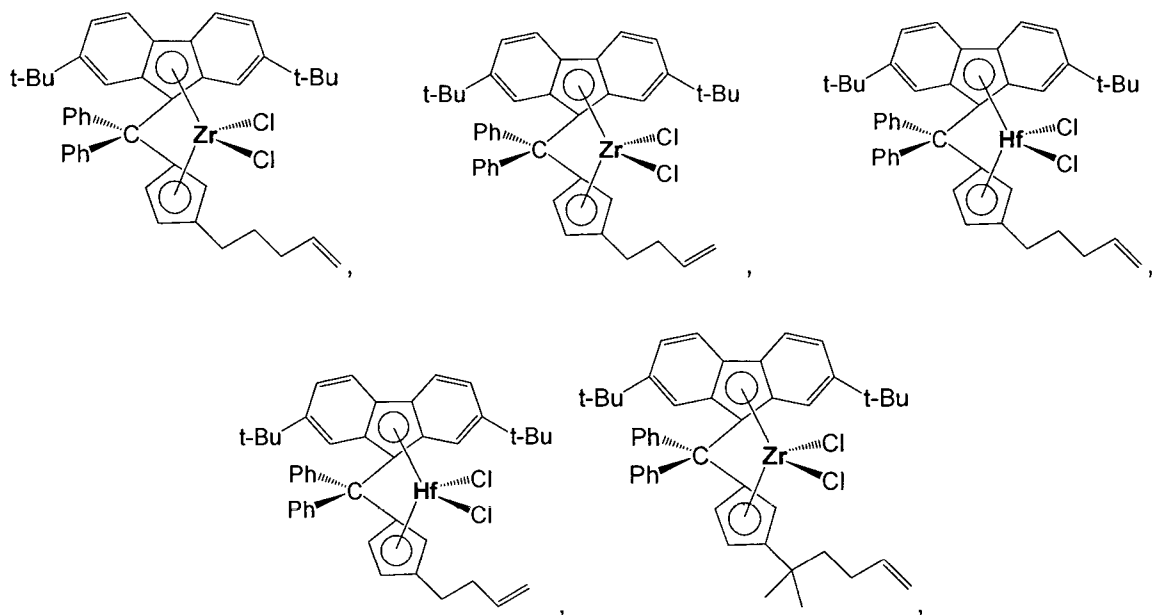
- M¹ is zirconium or hafnium;
 X is, independently, F, Cl, Br, or I; or
 X is, independently, H, BH₄, methyl, phenyl, benzyl, neopentyl, trimethylsilylmethyl, CH₂CMe₂Ph; CH₂SiMe₂Ph;
 CH₂CMe₂CH₂Ph; or CH₂SiMe₂CH₂Ph;
 E is C;
 R¹ and R² are independently an aryl group, either of which having up to 10 carbon atoms;
 R^{3A} and R^{3B} are independently a hydrocarbyl group or a trihydrocarbylsilyl group, any of which having up to 20
 carbon atoms; or hydrogen;
 n is an integer from 0 to 10, inclusive; and
 R^{4A} and R^{4B} are independently a hydrocarbyl group having up to 12 carbon atoms, or hydrogen.
2. The compound according to Claim 1, wherein X is independently, F, Cl, Br, or I.
3. The compound according to Claim 1, wherein X is Cl.
4. The compound according to Claim 1, wherein X is, independently, H, BH₄, methyl, phenyl, benzyl, neopentyl,
 trimethylsilylmethyl, CH₂CMe₂Ph; CH₂SiMe₂Ph; CH₂CMe₂CH₂Ph; or CH₂SiMe₂CH₂Ph
5. The compound according to Claim 1, wherein X is F, Cl, Br, or I, R^{3A} and R^{3B} are independently H, methyl, allyl,
 benzyl, butyl, pentyl, hexyl, or trimethylsilyl and n is an integer from 1 to 6, inclusive, and R^{4A} and R^{4B} are inde-
 pendently a hydrocarbyl group having up to 6 carbon atoms, or hydrogen.
6. The compound according to Claim 1, wherein X is Cl, Br, or I, R¹ and R² are phenyl, R^{3A} and R^{3B} are independently
 H or methyl, n is 1 or 2, and R^{4A} and R^{4B} are independently H or t-butyl.
7. The compound according to Claim 1 having the formula:





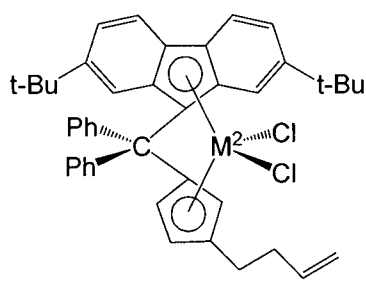
or any combination thereof.

8. The compound according to Claim 1 having the formula:



or any combination thereof.

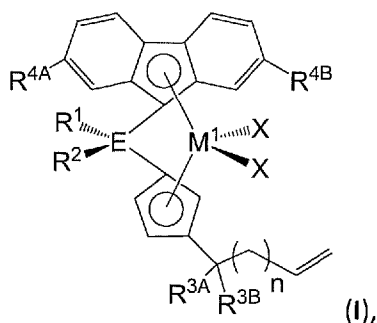
9. The compound of Claim 1 having the formula:



wherein M^2 is Zr or Hf.

Patentansprüche

1. Verbindung der Formel:



wobei

M^1 Zirkonium oder Hafnium ist;

X unabhängig F, Cl, Br, oder I ist; oder

X unabhängig H, BH_4 , Methyl, Phenyl, Benzyl, Neopentyl, Trimethylsilylmethyl, CH_2CMe_2Ph ; CH_2SiMe_2Ph ; $CH_2CMe_2CH_2Ph$; oder $CH_2SiMe_2CH_2Ph$ ist;

E C ist;

R^1 und R^2 unabhängig voneinander eine Arylgruppe sind, von denen jede bis zu 10 Kohlenstoffatome aufweisen kann;

R^{3A} und R^{3B} unabhängig voneinander eine Hydrocarbylgruppe oder eine Trihydrocarbysilylgruppe, von denen jede bis zu 20 Kohlenstoffatome aufweisen kann, oder Wasserstoff sind;

n eine ganze Zahl von einschließlich 0 bis 10 ist; und

R^{4A} und R^{4B} unabhängig voneinander eine Hydrocarbylgruppe mit bis zu 12 Kohlenstoffatomen oder Wasserstoff sind.

2. Verbindung nach Anspruch 1, wobei X unabhängig F, Cl, Br oder I ist.

3. Verbindung nach Anspruch 1, wobei X Cl ist.

4. Verbindung nach Anspruch 1, wobei X unabhängig H, BH_4 , Methyl, Phenyl, Benzyl, Neopentyl, Trimethylsilylmethyl, CH_2CMe_2Ph ; CH_2SiMe_2Ph ; $CH_2CMe_2CH_2Ph$; oder $CH_2SiMe_2CH_2Ph$ ist.

5. Verbindung nach Anspruch 1, wobei X F, Cl, Br oder I ist, R^{3A} und R^{3B} unabhängig voneinander H, Methyl, Allyl, Benzyl, Butyl, Pentyl, Hexyl oder Trimethylsilyl sind und n eine ganze Zahl einschließlich von 1 bis 6 ist, und R^{4A} und R^{4B} unabhängig voneinander einer Hydrocarbylgruppe mit bis zu 6 Kohlenstoffatomen oder Wasserstoff sind.

6. Verbindung nach Anspruch 1, wobei X Cl, Br oder I ist, R^{3A} und R^{3B} unabhängig voneinander H oder Methyl sind, n 1 oder 2 ist, und R^{4A} und R^{4B} unabhängig voneinander H oder t-Butyl sind.

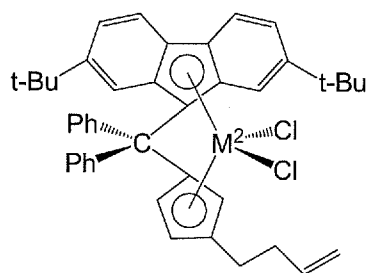


8. Verbindung nach Anspruch 1 der Formel:



oder jede beliebige Kombination davon.

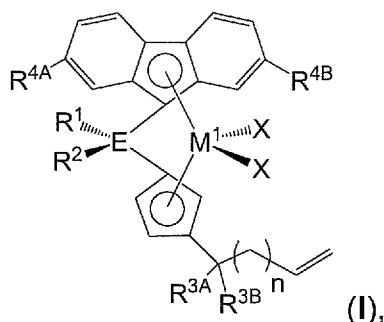
9. Verbindung nach Anspruch 1 der Formel:



wobei M^2 Zr oder Hf ist.

Revendications

1. Composé de formule :



dans lequel

M^1 représente zirconium ou hafnium ;

X représente, indépendamment, F, Cl, Br ou I ; ou

X représente, indépendamment, H, BH_4 , méthyle, phényle, benzyle, néopentyle, triméthylsilylméthyle, CH_2CMe_2Ph ; CH_2SiMe_2Ph ; $CH_2CMe_2CH_2Ph$; ou $CH_2SiMe_2CH_2Ph$;

E représente C ;

R^1 et R^2 représentent indépendamment un groupe aryle, l'un ou l'autre de ceux-ci possédant jusqu'à 10 atomes de carbone ; R^{3A} et R^{3B} représentent indépendamment un groupe hydrocarbyle ou un groupe trihydrocarbylsilyle, n'importe lequel de ceux-ci possédant jusqu'à 20 atomes de carbone ; ou l'hydrogène ;

n est un nombre entier allant de 0 à 10, inclus ; et

R^{4A} et R^{4B} représentent indépendamment un groupe hydrocarbyle possédant jusqu'à 12 atomes de carbone, ou l'hydrogène.

2. Composé selon la revendication 1, dans lequel X représente indépendamment F, Cl, Br ou I.

3. Composé selon la revendication 1, dans lequel X représente Cl.

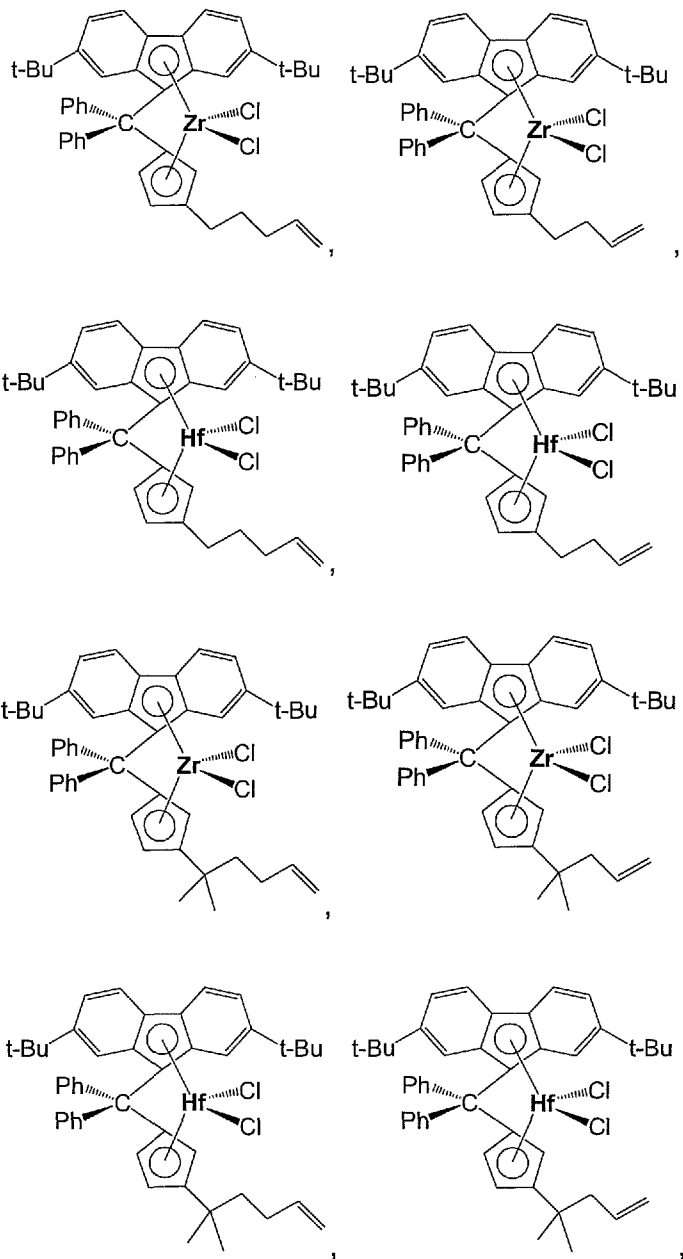
4. Composé selon la revendication 1, dans lequel X représente, indépendamment, H, BH_4 , méthyle, phényle, benzyle, néopentyle, triméthylsilylméthyle, CH_2CMe_2Ph ; CH_2SiMe_2Ph ; $CH_2CMe_2CH_2Ph$; ou $CH_2SiMe_2CH_2Ph$.

5. Composé selon la revendication 1, dans lequel X représente F, Cl, Br ou I, R^{3A} et R^{3B} représentent indépendamment H, méthyle, allyle, benzyle, butyle, pentyle, hexyle ou triméthylsilyle et n est un nombre entier allant de 1 à 6, inclus, et R^{4A} et R^{4B} représentent indépendamment un groupe hydrocarbyle possédant jusqu'à 6 atomes de carbone, ou

l'hydrogène.

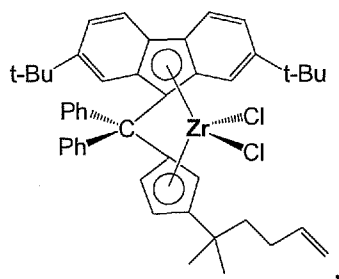
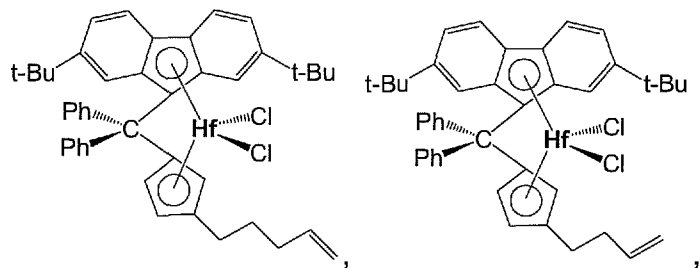
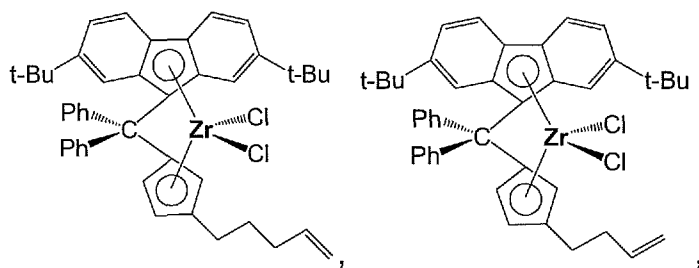
6. Composé selon la revendication 1, dans lequel X représente Cl, Br ou I, R¹ et R² représentent phényle, R^{3A} et R^{3B} représentent indépendamment H ou méthyle, n vaut 1 ou 2, et R^{4A} et R^{4B} représentent indépendamment H ou t-butyle.

7. Composé selon la revendication 1, de formule :



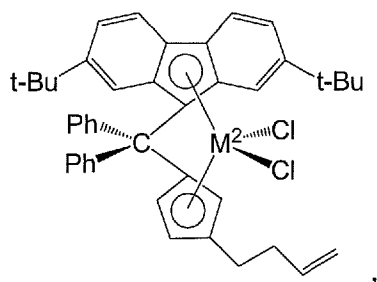
ou n'importe quelle combinaison de celles-ci.

8. Composé selon la revendication 1, de formule :

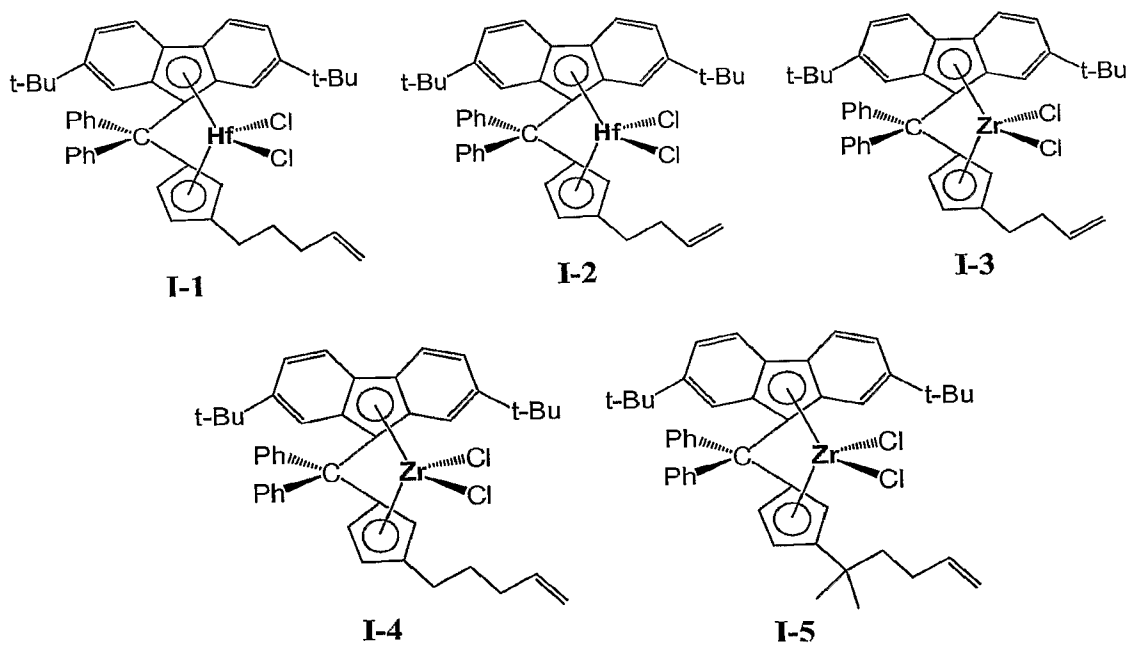


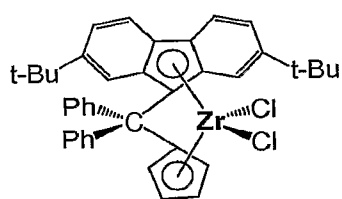
ou n'importe quelle combinaison de celles-ci ;

9. Composé selon la revendication 1, de formule

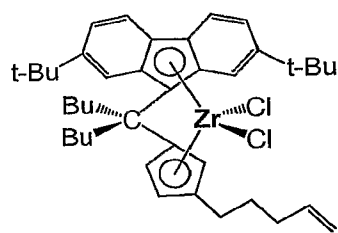


dans laquelle M² représente Zr ou Hf.

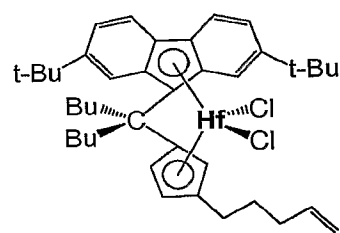




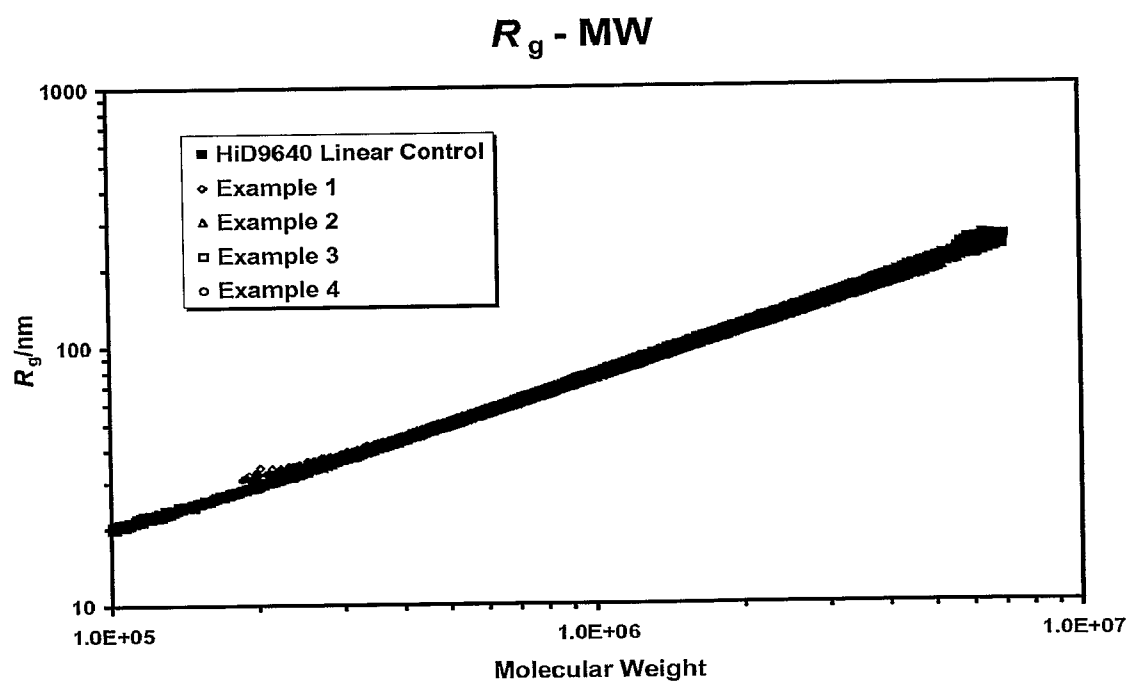
C-1

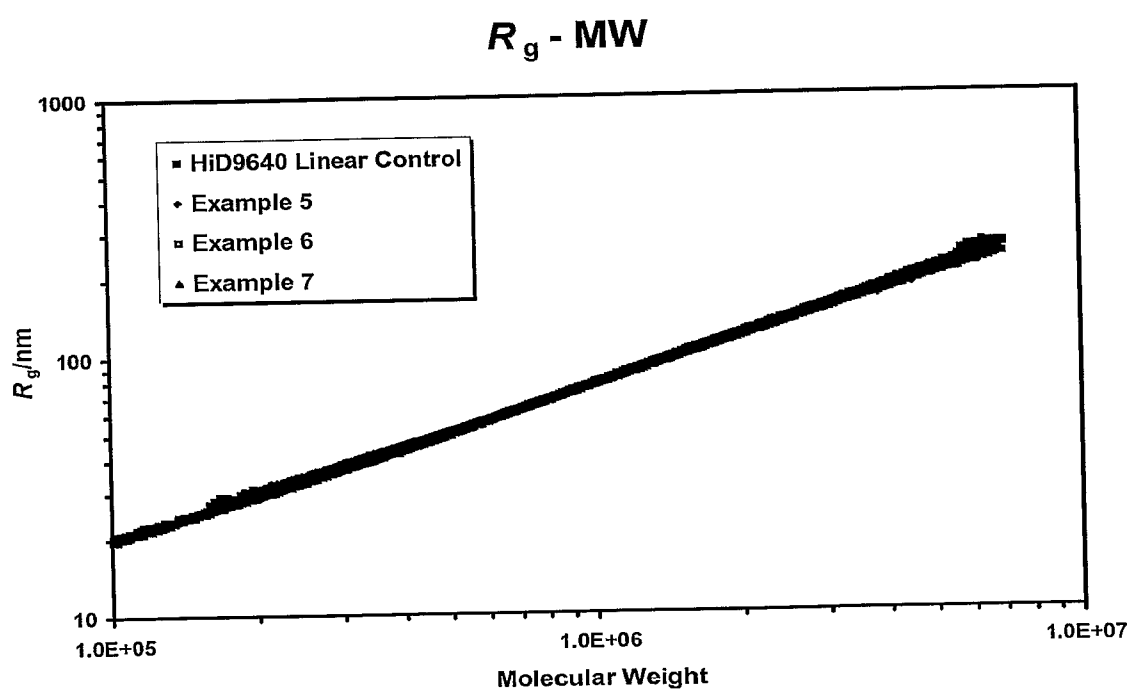


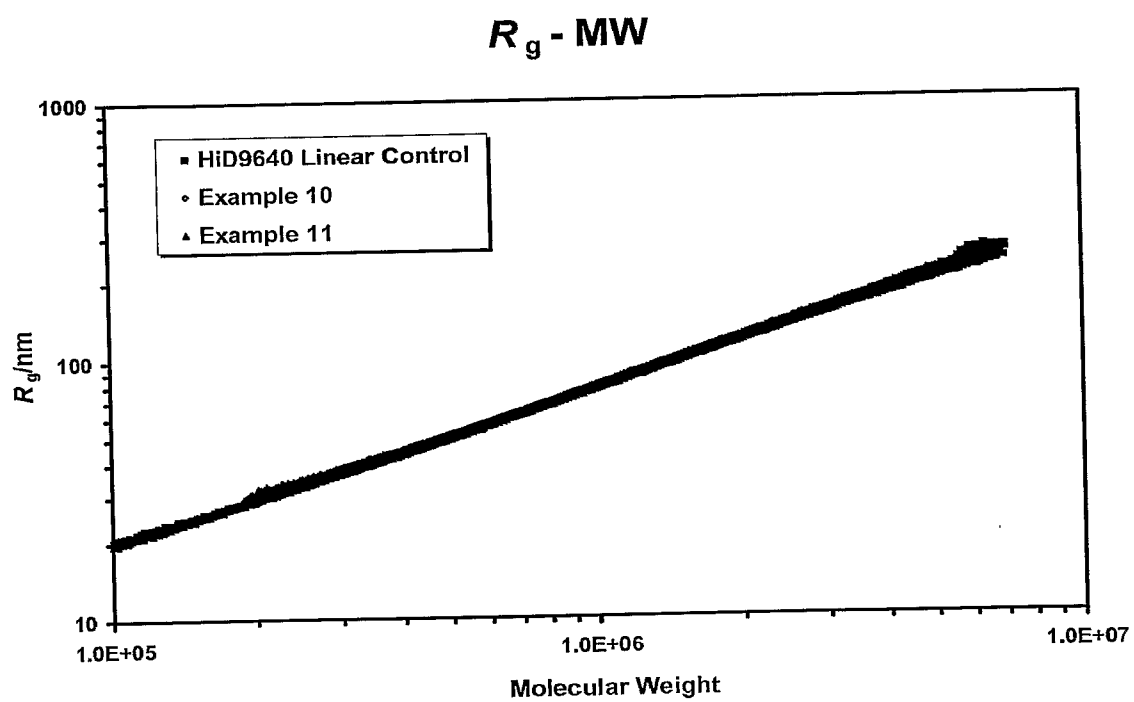
C-2

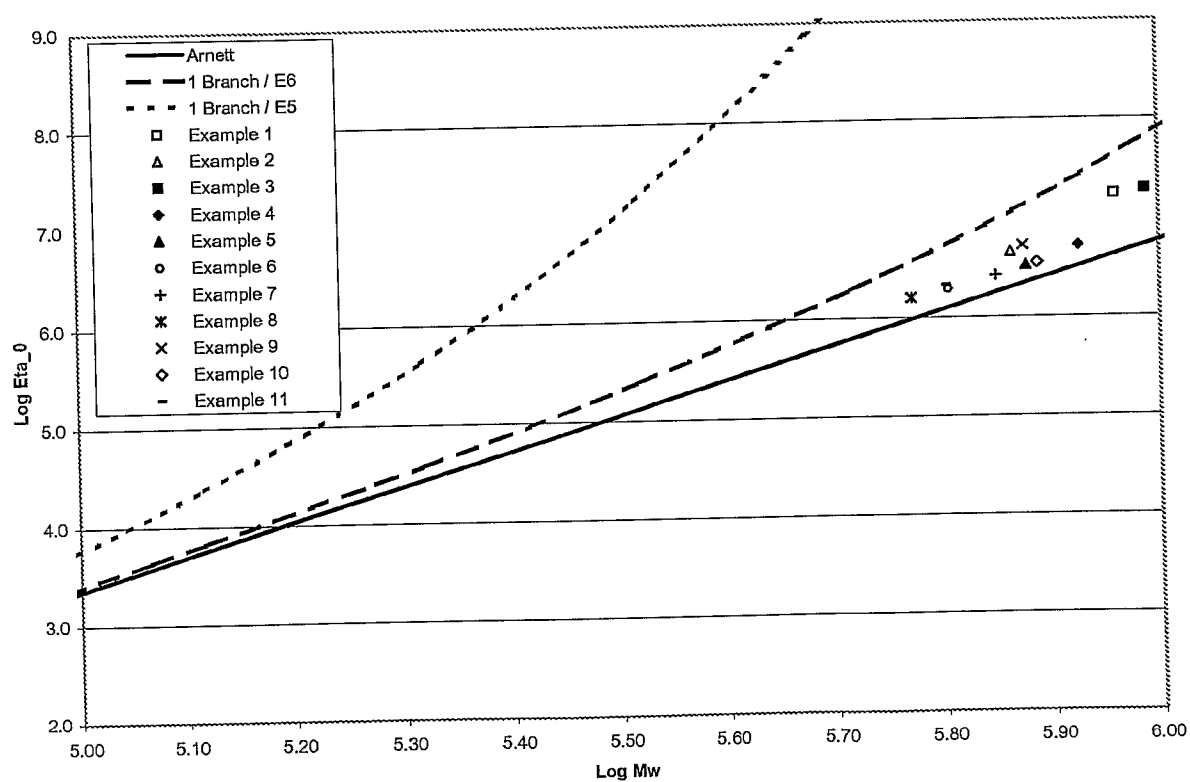


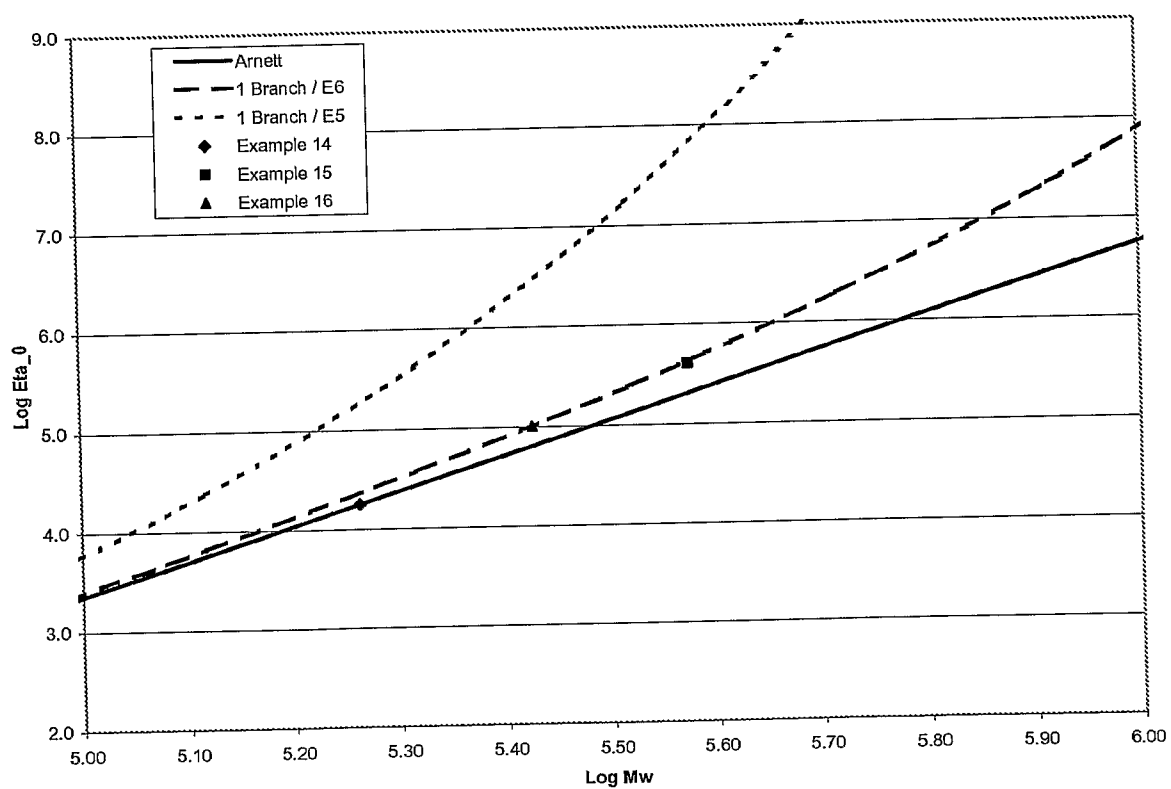
C-3

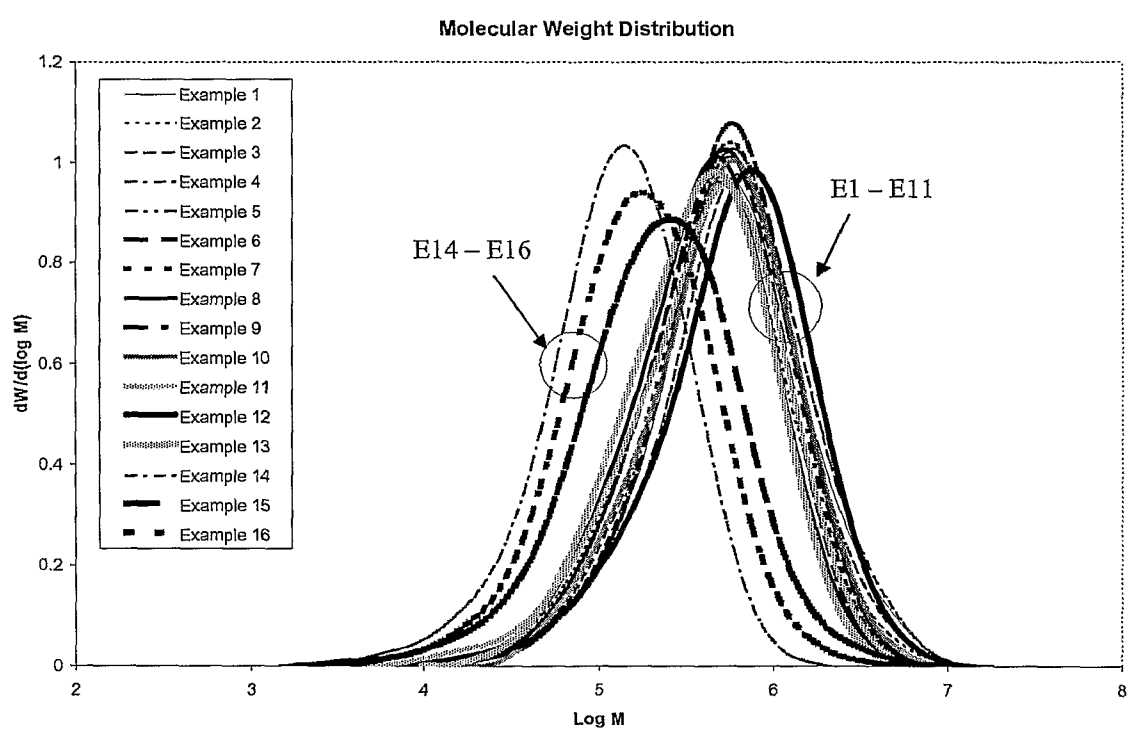












REFERENCES CITED IN THE DESCRIPTION

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Patent documents cited in the description

- US 5886202 A [0004]
- US 4939217 A [0016]
- US 5191132 A [0016]
- US 5210352 A [0016]
- US 5347026 A [0016]
- US 5399636 A [0016]
- US 5401817 A [0016]
- US 5420320 A [0016]
- US 5436305 A [0016]
- US 5451649 A [0016]
- US 5496781 A [0016]
- US 5498581 A [0016]
- US 5541272 A [0016]
- US 5554795 A [0016]
- US 5563284 A [0016]
- US 5565592 A [0016]
- US 5571880 A [0016]
- US 5594078 A [0016]
- US 5631203 A [0016]
- US 5631335 A [0016]
- US 5654454 A [0016]
- US 5668230 A [0016]
- US 5705578 A [0016]
- US 5705579 A [0016]
- US 6187880 B [0016]
- US 6509427 B [0016]
- US 6107230 A [0056]
- US 6165929 A [0056]
- US 6294494 B [0056]
- US 6300271 B [0056]
- US 6316553 B [0056]
- US 6355594 B [0056]
- US 6376415 B [0056]
- US 6391816 B [0056]
- US 6395666 B [0056]
- US 6524987 B [0056]
- US 6548441 B [0056]
- US 4452910 A [0060] [0154]
- US 5376611 A [0060] [0154]
- US 4060480 A [0060] [0154]
- US 4794096 A [0085]
- US 3242099 A [0088]
- US 4808561 A [0088]
- US 5919983 A [0093]
- US 5576259 A [0095]
- US 5807938 A [0095]
- US 3248179 A [0124]
- US 4501885 A [0124]
- US 5565175 A [0124]
- US 5575979 A [0124]
- US 6239235 B [0124]
- US 6262191 B [0124]
- US 6833415 B [0124]
- US 5455314 A [0125]
- US 5352749 A [0126]
- US 4588790 A [0126]
- US 5436304 A [0126]

Non-patent literature cited in the description

- KÖPPL, A. ; ALT, H. G. *J. Mol. Catal. A.*, 2001, vol. 165, 23 [0016]
- KAJIGAESHI, S. ; KADOWAKI, T. ; NISHIDA, A. ; FUJISAKI, S. *The Chemical Society of Japan*, 1986, vol. 59, 97 [0016]
- ALT, H. G. ; JUNG, M. ; KEHR, G. *J. Organomet. Chem.*, 1998, vol. 562, 153-181 [0016]
- ALT, H. G. ; JUNG, M. *J. Organomet. Chem.*, 1998, vol. 568, 87-112 [0016]
- JUNG, M. Doctoral Dissertation. University of Bayreuth, 1997 [0016]
- PIEFER, B. Doctoral Dissertation. University of Bayreuth, 1995 [0016]
- ZENK, R. Doctoral Dissertation. University of Bayreuth, 1994 [0016]
- WAILES, P. C. ; COUTTS, R. S. P. ; WEIGOLD, H. *Organometallic Chemistry of Titanium, Zirconium, and Hafnium*. Academic, 1974 [0016]
- CARDIN, D. J. ; LAPPERT, M. F. ; RASTON, C. L. *Chemistry of Organo-Zirconium and -Hafnium Compounds*. Halstead Press, 1986 [0016]
- Hawley's Condensed Chemical Dictionary. John Wiley & Sons, 1995 [0030]
- COTTON, F.A. ; WILKINSON, G. ; MURILLO, C. A. ; BOCHMANN, M. *Advanced Inorganic Chemistry*. Wiley-Interscience, 1999 [0030]
- T.J. PINNAVAIA. *Science*, 1983, vol. 220 (4595), 365-371 [0060] [0154]
- J.M. THOMAS. *Intercalation Chemistry*. Academic Press, Inc, 1972, 55-99 [0060]

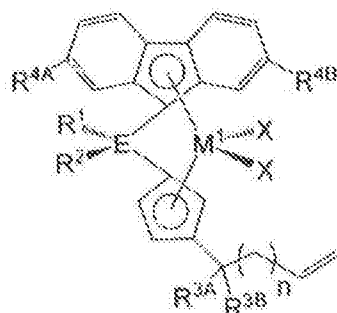
EP 1 984 408 B1

- Diagnosing long-chain branching in polyethylenes. *J. Mol. Struct.*, 1999, vol. 485-486, 569-584 [0137]
- *J. Phys. Chem.*, 1980, vol. 84, 649 [0138]
- **J.M. THOMAS**. Intercalation Chemistry. Academic Press, Inc, 1972, vol. 3, 55-99 [0154]
- Dynamics of Polymeric Liquids. **R. BYRON BIRD ; ROBERT C. ARMSTRONG ; OLE HASSAGER**. Fluid Mechanics. John Wiley & Sons, 1987, vol. 1, 171-172 [0168]
- **WYATT, P.J.** *Anal. Chim. Acta*, 1993, vol. 272, 1 [0169]
- **K. J. STONE ; R. D. LITTLE**. *J. Org. Chem.*, 1984, vol. 49 (11), 1849-1853 [0183]
- **K.J. STONE ; R.D. LITTLE**. *J. Org. Chem.*, 1984, vol. 49 (11), 1849-1853 [0187]

SZABADALMI IGÉNYPONTOK



1. Az alábbi általános képletű vegyület



(I), ahol

M^1 jelentése cirkónium vagy hafnium atom;

X jelentése, függetlenül, Cl, Br vagy I atom; vagy

X jelentése, függetlenül, H atom, BH_4 , metil-, fenil-, benzil-, neopentil-, trimetil-szililmetil-csoport, CH_2CMe_2Ph ; CH_2SiMe_2Ph ; $CH_2CMe_2CH_2Ph$; vagy $CH_2SiMe_2CH_2Ph$;

E jelentése C atom;

R^1 és R^2 függetlenül arilcsoportot képvisel, melyek bármelyike legfeljebb 10 szénatomos;

R^{3A} és R^{3B} függetlenül szénhidrogén-csoport vagy trihidrokarbil szilil-csoport, melyek bármelyike legfeljebb 20 szénatomos; vagy hidrogénatom;

n értéke 0-10, beleértve a szélső értékeket; és

R^{4A} és R^{4B} függetlenül legfeljebb 12 szénatomos szénhidrogén csoport, vagy hidrogénatom.

2. Az 1. igénypont szerinti vegyület, ahol X jelentése függetlenül F, Cl, Br vagy I atom.

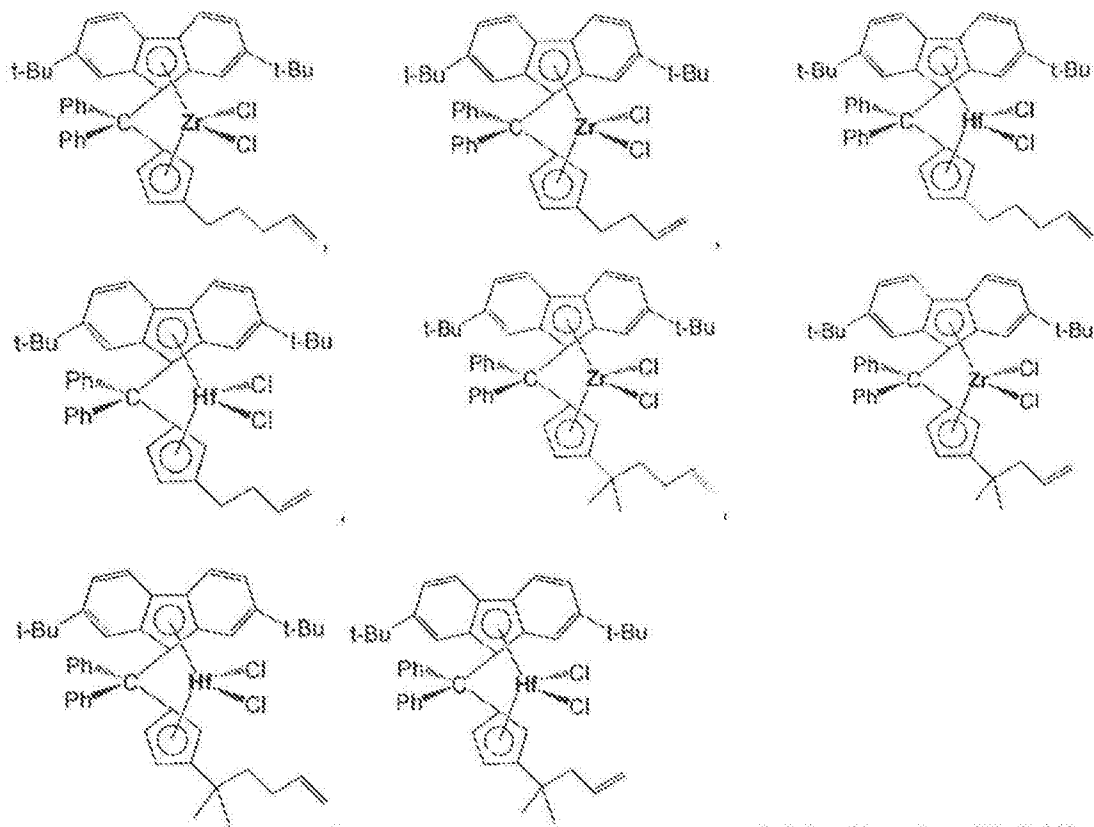
3. Az 1. igénypont szerinti vegyület, ahol X jelentése Cl atom.

4. Az 1. igénypont szerinti vegyület, ahol X jelentése, függetlenül, H atom, BH_4 , metil-, fenil-, benzil-, neopentil-, trimetil-szililmetil-, CH_2CMe_2Ph ; CH_2SiMe_2Ph ; $CH_2CMe_2CH_2Ph$; vagy $CH_2SiMe_2CH_2Ph$ csoport.

5. Az 1. igénypont szerinti vegyület, ahol X jelentése F, Cl, Br vagy I atom, R^{3A} és R^{3B} függetlenül H atom, metil-, allil-, benzil-, butil-, pentil-, hexil- vagy trimetilszilil-csoport és n értéke 1-6, beleértve a szélső értékeket, és R^{4A} és R^{4B} függetlenül legfeljebb 6 szénatomos szénhidrogén csoport vagy hidrogénatom.

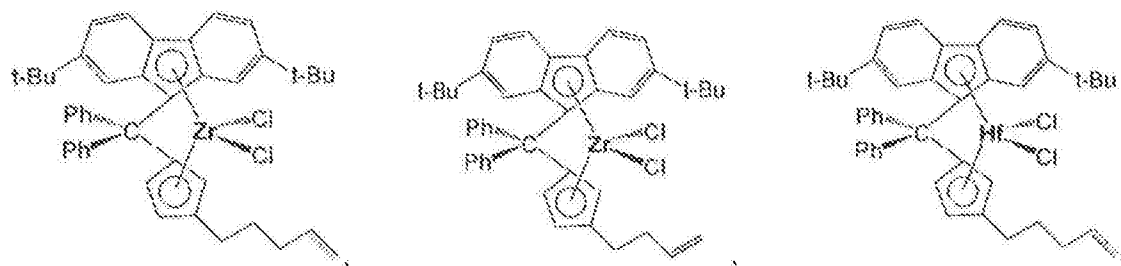
6. Az 1. igénypont szerinti vegyület, ahol X jelentése Cl, Br vagy I atom, R^1 és R^2 fenilcsoport, R^{3A} és R^{3B} függetlenül H atom vagy metilcsoport, n értéke 1 vagy 2, és R^{4A} és R^{4B} függetlenül H atomot vagy terc-butilcsoportot képvisel.

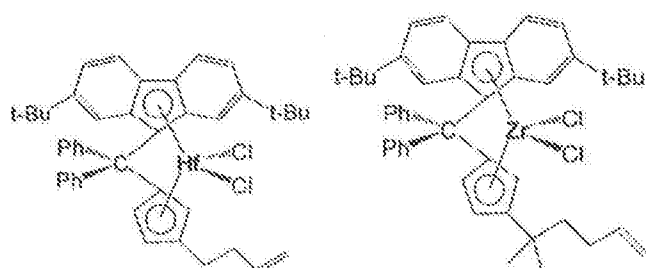
7. Az 1. igénypont szerinti vegyület, mely az alábbi képlettel rendelkezik:



vagy ezek bármilyen kombinációja.

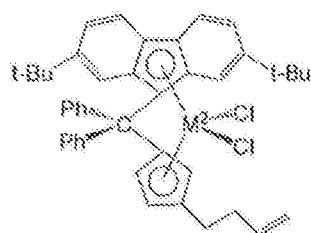
8. Az 1. igénypont szerinti vegyület, mely az alábbi képlettel rendelkezik:





vagy ezek bármilyen kombinációja.

9. Az 1. igénypont szerinti vegyület, mely az alábbi képlettel rendelkezik



ahol M^2 jelentése Zr vagy Hf atom.