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(54) Title: HYDROMETALLURGICAL PROCESS AND APPARATUS FOR RECOVERING METALS FROM WASTE MATERIAL

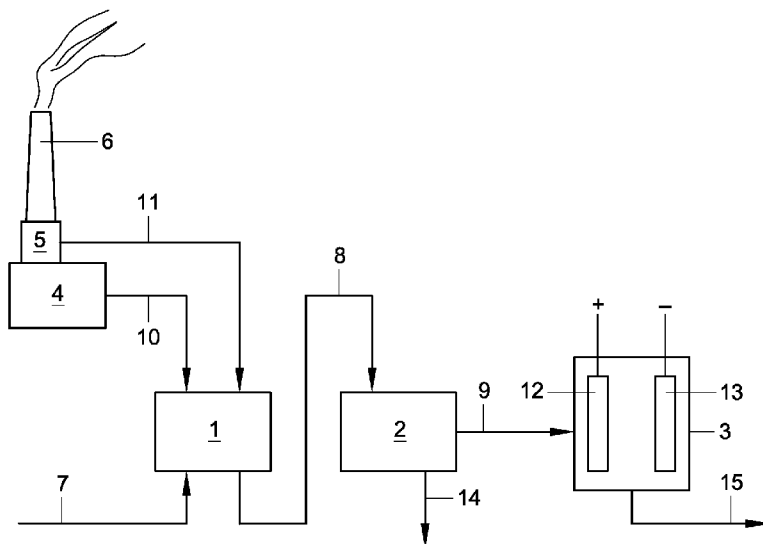


Fig. 2

(57) Abstract: The invention is directed to a process and apparatus for recovering metals, in particular for recovering metals from bottom ash from incineration plants, such as municipal waste incineration plants (4). In accordance with the invention, a feed containing ash is fed to an oxidizing unit (1), wherein at least part of said metals is oxidized in the presence of one or more acids and at least one oxygen donor, thus producing a stream comprising metal ions. From this stream the metals of interest are selected and concentrated in a solvent extraction unit (2) and thereafter converted into metallic form in an electrowinning unit (3).

WO 2011/074948 A1

## HYDROMETALURGICAL PROCESS AND APPARATUS FOR RECOVERING METALS FROM WASTE MATERIAL

The invention is directed to a process and apparatus for recovering metals, in particular for recovering metals from ash from incineration plants, such as municipal waste incineration plants.

Incineration of waste materials generally produces bottom ash and  
5 flue gases, which may contain particulates. These particulates are called flue dust or fly ashes. These fly ashes are mainly composed of calcium oxide, silicon dioxide and various metal oxides. Normally, the flue gases are subjected to a cleaning step prior to venting into the atmosphere. The bottom ash contains all non-combustible material, typically salts and metals. The metals may be  
10 present in their neutral (metallic) form and/or as salts, in which the metals are present in the form of ions. Its composition depends strongly on the feed of the incineration process.

In the art the bottom ash from incineration plants is usually disposed of by dumping or using the material for construction purposes after  
15 immobilizing the leachable compounds, *e.g.* by packing the ash in bentonite.

The availability of metals such as copper has been dropping in recent years as a result of rising demand and the exhaustion of natural resources.

Present methods recover valuable metals such as copper from waste  
20 materials using extraction methods based on a combination of magnetic and mechanical extraction. These known methods are not very efficient for removing particles having sizes below 4 mm and are considered infeasible for particle sizes below 1 mm. Furthermore, these magnetic/mechanical methods generally produce scrap metal mixtures, which need to be further refined  
25 before they can be reused. Moreover, these methods omit the metal salts, which are the main cause of environmental issues, such as ground water pollution, when above mentioned ashes are used for construction purposes.

EP-A-0 274 059 describes a process for the recycling metals from pyrolysed electrical batteries, assembled print circuit boards and electronic

components using electrolysis. The metals collected during the electrolysis step described in EP-A-0 274 059 are then separated from each other following the electrolysis step using known metallurgical or electrochemical methods.

However this process has the disadvantage that a further processing step is  
5 required to separate the metals following electrolysis. Also, the process of EP-A-0 274 059 is limited in its use to electrically conductive input streams. The present invention seeks *inter alia* to provide a method that is more versatile so that it may be used also for processing streams that are not, or only slightly electrically conductive. This is in particular the case for bottom ash, which is  
10 not electrically conductive.

The present invention seeks to overcome some of the disadvantages of processes currently used and provide a process and an apparatus to recover valuable metals, including but not limited to copper, tin, gold, silver, zinc, lead, platinum, and mixtures thereof, from sources of low or no non electrical  
15 conductivity, such as bottom ash streams produced by incineration processes.

It was found that this object can be met by providing a process for recovering one or more metallic metals from a waste stream of ash containing said metals, which process comprises:

- feeding said waste stream of ash to a leaching or oxidizing unit,  
20 wherein at least part of said metallic metals is oxidized in the presence of one or more acidic substances and at least one oxygen donor, thus producing a stream comprising the corresponding metal ions;

- concentrating one or more metal ions from said stream comprising them, thus producing a stream of concentrated metal ions;

25 - reducing at least part of the metal ions from said stream of concentrated metal ions, thus producing said one or more of metals in metallic form.

The waste supply stream of ash preferably is the bottom ash from an incineration plant and more preferably an incineration plant for municipal  
30 waste. The process of the invention is particularly suitable for treating bottom

ash because bottom ash contains economically significant levels of metals and is produced in large quantities.

The metals of interest are present in the ash that is used as a feed for the present invention in their metallic, their oxidized form or other salt form. Usually, the ash contains several different types of metals, *e.g.* mixtures of copper and other metals. The present inventors realized that typically the metals of interest, in particular copper, are frequently present in a metallic form, *viz.* not as a salt or as ions, and that in order to select this metal and concentrate it, it is needed to carry out an oxidation step first, which is most effectively carried out in the presence of an acid and oxygen, since this greatly accelerates the oxidation reaction. The process carried out in the oxidation step is also known as leaching. Leaching is generally used to dissolve metal salts in aqueous media. Leaching is advantageous because it is relatively easy to perform and gaseous pollution is minimized. Although the leaching is said to have a low efficiency, the present inventors realized that this is not prohibitive for the object of the present invention.

In the art, leaching is usually performed on sulfide or oxide ores, with the notable exception of gold. Ores containing metallic or native metals are usually treated with gravity methods or the aforementioned mechanical methods, as they cannot be significantly leached with the traditional methods within appreciable time. The present invention is therefore characterized in that it treats waste streams containing metallic metals.

Unlike some existing hydrometallurgical methods which leach metallic metals from waste streams, such as a recovery process of zinc from the thin zinc coating on galvanized steel like US-B-6 878 356, the present invention accelerates the oxidation of the metallic metal with an acidic substance and an oxygen donor, thereby making it feasible to dissolve not only thin coatings, but also larger pieces of metal.

The first step, where the metal is oxidized, can take place under acidic conditions or for some metals also under alkaline conditions, as the term

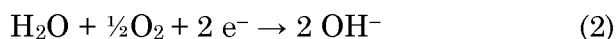
acidic substance as used herein refers to either a substance that dissociates in water to yield hydrogen ions ( $H^+$ ) or a substance that can accept a pair or electrons, *i.e.* a Lewis acid. The half reaction here is for example for hydrogen ions:

5



Under alkaline aqueous conditions, the equivalent half reaction is sometimes written as follows:

10



The half reaction (1) is identical to (2) when to the left and right of the equation, one water molecule ( $H_2O$ ) is removed and two hydroxide ions ( $OH^-$ ) are added.

15

For the leach step to be feasible under alkaline conditions, the metal ions need to be dissolved in the alkaline solution in significant quantities. To be economically feasible, currently at least 1 to 5 g/L is needed for copper. To dissolve the metal ions in an alkaline solution, forming metal complexes is a suitable approach.

20

For some metals, such as zinc or lead, the hydroxide ions themselves can form complexes with the metal ions, for example producing the water soluble zincate. Currently, for an economical process the concentration of hydroxide ions needs to be very high to dissolve sufficient metal ions, *i.e.* 5-10 molar for zinc. The alkaline solution can be obtained by adding a basic substance which may comprise soluble salts, such as sodium or potassium hydroxide.

25

For transition metal ions, such as copper, zinc and silver, ammonia can be used to complex the metal ions and make them more soluble. For other metals, other water soluble chelating agents can be used to complex the metal

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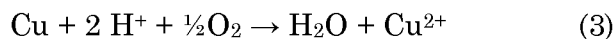
ions, examples are EDTA and D<sub>2</sub>EHPA (di(2-ethylhexyl) phosphate acid). The use of ammonia or other chelating agents is more advantageous than using hydroxide to complex the metal ions, as the concentration of hydroxide ions need not be so high, and as the solvent is water, therefore the concentration of hydrogen ions can be higher. This increases the reaction rate of half reaction (1), as more reactant hydrogen ions are available. To increase the concentration of these hydrogen ions, an acid can be added to the solution, such as hydrochloric acid, dissolved carbon dioxide, or sulfuric acid.

When the feed stream has a high carbonate salt contents, such as bottom ash, and the required product is copper or zinc, using a combination of ammonia solution and dissolved carbon dioxide is highly advantageous. The dissolved carbon dioxide forms carbonic acid and with the solid carbonate salts, the solution is pH buffered around a pH of 9. This way, the large emissions of carbon dioxide, which occurs at lower pH, is avoided. Furthermore, about 50% of the dissolved ammonia is protonated at a pH of 9 to form ammonium ions: a very suitable hydrogen ion source for half reaction (1). The remainder dissolved ammonia can then be used to complex the copper or zinc ions. Thirdly, the ammonia selectively complexes the required product, while the dissolved carbon dioxide precipitates with unwanted ions. This reduces waste water significantly. And fourthly, carbon dioxide is amply available at incineration processes.

With great advantage, the acid is obtained from an off-gas treatment step, *viz.* a step in which harmful components, such as NO<sub>x</sub>, chlorine, metal oxides, and sulfur oxides (notably SO<sub>2</sub> and SO<sub>3</sub>) are removed from the off-gas of for example a waste incinerator. In these known off-gas treatment steps, the polluted gas stream is typically contacted with a liquid, for instance by scrubbing the gas with a liquid in a wet scrubber, by spraying the liquid through the gas, by forcing the gas through a pool of liquid, or by some other contact method. The acids are created in the process of neutralizing the corrosive gases. For example, when chlorine containing products, such as

polyvinyl chloride (PVC) are incinerated, HCl gas is formed, which creates the hydrochloric acid when in contact with water. These off-gas treatment operations commonly result in streams of acid, preferably strong acids, such as sulfuric acid, hydrochloric acid and combinations thereof. Since normally the  
5 off gas treatment is carried out on site, directly after the incineration has taken place, these acids streams can be put to use with great economic advantage. An added advantage is that they do not have to be treated in waste water treatment plant at great costs.

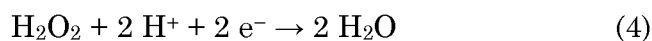
As mentioned above, it was further found that applying oxygen in  
10 the form of air helps to accelerate the oxidation reaction of the metallic metals. Typically the following reaction occurs, wherein copper is used as an example:



15 Additionally, blowing the air in the unit provides for an excellent mixing in the oxidation step, which is carried out in the liquid phase. By allowing air to flow through the oxidizing unit, the oxidation step can thus be considerably sped up.

Alternatively, or in addition, a liquid oxidizing donor such as  
20 hydrogen peroxide may be used. Any copper in the supply stream acts as a catalyst, liberating the oxygen from the hydrogen peroxide. This way, the availability of oxygen increases and the reaction rate is increased. The hydrogen peroxide itself can also act as an oxidizing agent, with the following reduction reaction:

25



However, contrarily to what may be expected, reaction (4) does not promote the leaching of metal ions like copper or nickel. Instead, a protecting passive  
30 layer is formed on the surface of those metals, preventing further significant

leaching of them. Therefore, it is highly desirable to follow a correct order of adding reagents when using hydrogen peroxide as oxidizing donor. To this end, first the waste stream is immersed in a solution containing the acid or chelating agent, for example ammonia. Only then the hydrogen peroxide is  
5 added.

The leaching or oxidation is carried out in an aqueous medium at suitable conditions. The leaching progresses faster at higher temperatures, but due to energy usage considerations the temperature is preferably 20 to 30 °C. When surplus heat of the incinerator is available, the preferred temperature is  
10 60 to 80 °C. The pH of the leaching solution when not using a chelating agent is preferably between -1 to 4, more preferably from -1 to 1. The pH of the alkaline solution is preferably about 8-10, more preferably 9, which ensures that all the carbonates from the bottom ash remain in a solid form, while ensuring sufficient hydrogen ions are available for the reaction.

15 This results in high metal concentrations, for example in copper concentrations of 1 – 10 g/L.

After the metallic metals are at least partly oxidized, preferably to a considerable extent, *e.g.* 99% (based on weight) or more, the desired oxidized metal needs to be selected and concentrated. This is preferably done in a  
20 separate unit, in particular in a solvent extraction unit. It is also possible to feed the effluent directly to the electrowinning step and choose conditions in the electrowinning step such that a selection and concentration of the desired metal is obtained, either in oxidized form or in metallic form; however, in practice this is usually not easily accomplished. Solvent extraction is a liquid-  
25 liquid extraction and is a method to separate compounds based on their relative solubilities in two different liquids, which liquids are immiscible or poorly miscible with each other. Usually water and an organic solvent are used as the liquids. Usually a complexing agent is used to selectively extract the desired metal from solution that is obtained after the leaching step. The metal



can be recovered from it in a concentrated form after leaching. The result is an extraction of one substance from one liquid phase into the other liquid phase.

A very suitable extractant for this purpose is LIX 84<sup>TM</sup>, which is commercially obtainable from Cognis corporation. The desired metal is  
5 extracted from the pregnant leaching liquid into the organic extractant by mixing the two liquids and, after settling, the two liquids are separated by gravity. Next, the metal is stripped from the extractant with a concentrated acid against the concentration gradient, using a similar process of mixing and settling. This mixing and settling can, for example, be performed in a multiple  
10 stage mixer-settler setup, pulsed column, or other apparatus.

In the final step of the process of the present invention, the metal ions are reduced to form metal deposits in metallic form. Preferably this is done by subjecting a concentrated effluent containing substantially one single metal in oxidized form, *e.g.* Cu<sup>2+</sup> ions to the reduction step. Preferably this step  
15 is carried out using a galvanic cell, wherein a pair of a negative cathode and a positive anode is present. The metal ions are reduced on the cathode, thus forming a metal deposit thereon, which can later be easily removed. In one embodiment, the electrowinning step is carried out using the process and device described in WO-A-98/58090.

20 WO-A-98/58090 describes a method and device for recovering metals from a flow using an electrochemical cell. WO-A-98/58090 also does not describe or suggest any method of selecting ions of interest, nor that this may be achieved by means of solvent extraction.

In a preferred embodiment, the leaching (*viz.* the oxidization of the  
25 metallic metal into their corresponding ions) is improved further by applying an external voltage to the metal grains in the waste stream, as schematically depicted in figure 1. This is particularly useful to plate at least part of the copper present in the solution. The remainder of the metals can be treated with the separate solvent extraction and electrowinning steps. To apply the  
30 voltage, the ash is spread on a flat-shaped, nearly horizontal (for instance

having an angle of inclination with the surface of less than 10°, preferably between 4°-8°) inert anode, such that the metal grains touch the anode plate. The anode plate needs to be somewhat horizontal to support the ash. A small incline is useful to transport the ashes, *e.g.* after shaking the ashes move. The

5 anode can be shaped, such that the contact area with the metal grains is maximized, *e.g.* by providing it in the form of a gauze structure. It may for instance be made from a titanium alloy, *e.g.* ASTM grade 7 titanium. This anode is placed together with a cathode in a vessel filled with an acid and additives, for example sulfuric acid with dissolved copper sulfate and small

10 amounts of chloride and other additives. The cathode, for example made from Type 316 stainless steel, is located at a short distance, typically several centimeters, *e.g.* from 1-5 cm, above the anode. A controlled voltage difference (preferably between 0.1 to 1 V, *e.g.* around 0.3 V) is applied on the two electrodes, which will result in fast leaching (dissolving) of the metallic metals,

15 *viz.* oxidizing them into their corresponding ions. At the cathode the copper ions will directly plate. This embodiment is highly advantageous, because of the simplicity of the set up. In fact, in this embodiment, the three steps of oxidation, selecting and plating the ions to their corresponding metals are combined in a single piece of equipment.

20 The concentration of copper ions needs to be kept constant or substantially constant in the above setup. However, copper ions are expended for every non-copper metal that is leached, the leaching of iron follows the reaction:  $\text{Fe(s)} + \text{Cu}^{2+} \rightarrow \text{Fe}^{2+} + \text{Cu(s)}$ , which depletes the electrolyte of Cu.

In more detail, the following two half-reactions take place for iron

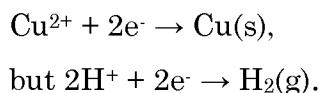
25 and copper:

at the anode  $\text{Fe(s)} \rightarrow \text{Fe}^{2+} + 2\text{e}^-$ ,

and at the cathode:  $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu(s)}$ .

For this reason, in one embodiment, first the less noble metals in a leach fluid are leached, optionally also accelerated by the setup where the

ashes are spread on the anode and cathode, but without dissolved copper sulfate. In that case the half-reaction at the cathode is not



5           The metallic metals that oxidize can be treated with the subsequent solvent extraction and electrowinning steps. When the applied electrical potential is sufficiently low (less than 0.34 V), half noble metals, such as copper, and noble metals, such as silver and gold, remain in solid form, and can therefore be treated in a leaching stage with copper sulfate according to  
10 the above mentioned method.

The acid that is expended can be obtained from wet gas scrubbers.

Figure 2 is a schematic representation of one embodiment that is suitable for carrying out the present invention. In figure 2, a waste stream of ash (10) that is produced in incinerator (4) is fed to oxidizing unit (1), wherein  
15 metal atoms are oxidized to their corresponding ions by feeding oxygen rich feed (7) and acid feed (11). In figure 2, acid feed (11) originates from the gas effluent treatment step (5) of the incinerator process, while the cleaned flue gas of the incinerator process is vented through chimney (6). The effluent (8) of oxidizer (1) is fed to solvent extraction unit (2), which is an example of carrying  
20 out the required selection and concentration of the desired metal ions. Effluent (9), containing a high concentration of ions of the desired metals is fed to electrowinning step (3). Effluent (14) of the solvent extraction step can be subjected to a separate extraction step to extract other ions for further processing. In electrowinning step (3) an anode (12) and a cathode (13) are  
25 used to reduce the metal ions into their metallic form. Effluent (15), now containing a lower concentration of ions, is returned to solvent extraction unit (2) for reuse.

#### Example

Two different experiments were performed on waste incinerator  
30 bottom ash, with a dry weight of  $150 \pm 5$  kg. 150 kg of 10w% ammonia dissolved

in water was added in a column with a diameter of 45 cm and a height of 2 m. In these experiments no acid was added, which resulted in an increase of pH due to acid consumption. In the first experiment air was bubbled through from the bottom of the column with a KNF Laboport™ 60 W diaphragm pump. In the second experiment the leach fluid was pumped through the column with two Prominent™ diaphragm pumps and hydrogen peroxide was continuously added to the recycled leach fluid. In the table below the amounts of leached metals are displayed.

Experiment	Day	[Cu] (mg/kg) *	[Zn] (mg/kg) *
10w% NH <sub>3</sub> + Air	6	1329	812
10w% NH <sub>3</sub> + Air	13	2356	1096
10w% NH <sub>3</sub> + Air	17	2882	1205
10w% NH <sub>3</sub> + H <sub>2</sub> O <sub>2</sub>	6	1612	783
10w% NH <sub>3</sub> + H <sub>2</sub> O <sub>2</sub>	13	2780	1082

\*) Based on dry matter

10

For these experiments the solvent extraction step was not included. When the solvent extraction was included, a faster leaching was observed.

In a separate setup the pregnant leach fluid was fed through the solvent extraction unit and a rich copper sulfate electrolyte containing 18560 ppm copper was successfully produced from dilute sulfuric acid. From this electrolyte a cathode was plated with copper.

15

Claims

1. Process for recovering one or more metallic metals from a waste supply stream containing them, which process comprises:
  - feeding said stream to an oxidizing unit, wherein at least part of said metallic metals is oxidized in the presence of one or more acidic substances, at least one oxygen donor and optionally a complexing agent, thus producing a stream
- 5 comprising metal ions; followed by a step of selecting the ions of interest and converting them into their metallic form.
2. Process according to claim 1, wherein said step of selecting the ions of interest and converting them into their metallic form comprises:
  - 10 - concentrating one or more of said metal ions from said stream comprising them, thus producing a stream of concentrated metal ions;
  - reducing at least part of the metal ions from said stream of concentrated metal ions, thus producing said one or more of metals in metallic form.
3. Process according to claim 1 or 2, wherein said oxygen donor is an
- 15 oxygen containing dissolved gas, preferably air, or a liquid, preferably hydrogen peroxide.
4. Process according to any of the previous claims, wherein the oxidation of said metallic metals is accelerated by an induced electrical potential by which the metals are made anodic.
- 20 5. Process according to any of the previous claims, wherein part of the metals is reduced on a cathode in the oxidizing unit, thus producing said metals in metallic form.
6. Process according to any of the previous claims, wherein said one or more acidic substances are selected from H<sub>2</sub>SO<sub>4</sub>, HCl, HNO<sub>3</sub>, CO<sub>2</sub> or
- 25 combinations thereof.

7. Process according to any of the previous claims, wherein one or more (complex forming) basic substances are selected from sodium hydroxide, potassium hydroxide or ammonia.
8. Process according to any of the previous claims, wherein said step of  
5 concentrating one or more metal ions is carried out in a solvent extraction step.
9. Process according to any of the previous claims, wherein said step of reducing at least part of the metal ions from said stream of concentrated metal ions is carried out in an electrowinning step.
- 10 10. Process according to any of the previous claims, wherein said metallic metal is copper.
11. Process according to any of the previous claims, wherein said waste supply stream containing said one or more metallic metals comprises ash, in particular bottom ash from an incineration plant, in particular an incineration  
15 plant for municipal waste.
12. Process according to any of the previous claims, wherein said one or more acidic substances are obtained from the flue gas treatment step of the incineration plant that produced the bottom ashes.
13. Apparatus for carrying out a process according to any of the previous  
20 claims, comprising:
- an oxidizing unit, which comprises an opening for feeding a feed comprising a waste stream containing metallic metals, in particular bottom ash from an incineration plant, a duct for feeding acid and a duct for feeding oxygen containing gas;
  - 25 - a solvent extraction unit, which is fitted with a duct for receiving the effluent of said oxidizing unit;
  - an electrowinning unit which is fitted with a duct for receiving the effluent of said oxidizing unit.

14. Apparatus according to claim 13, wherein said oxidizing unit further comprises a pair of electrodes, which preferably comprises a nearly horizontal anode.

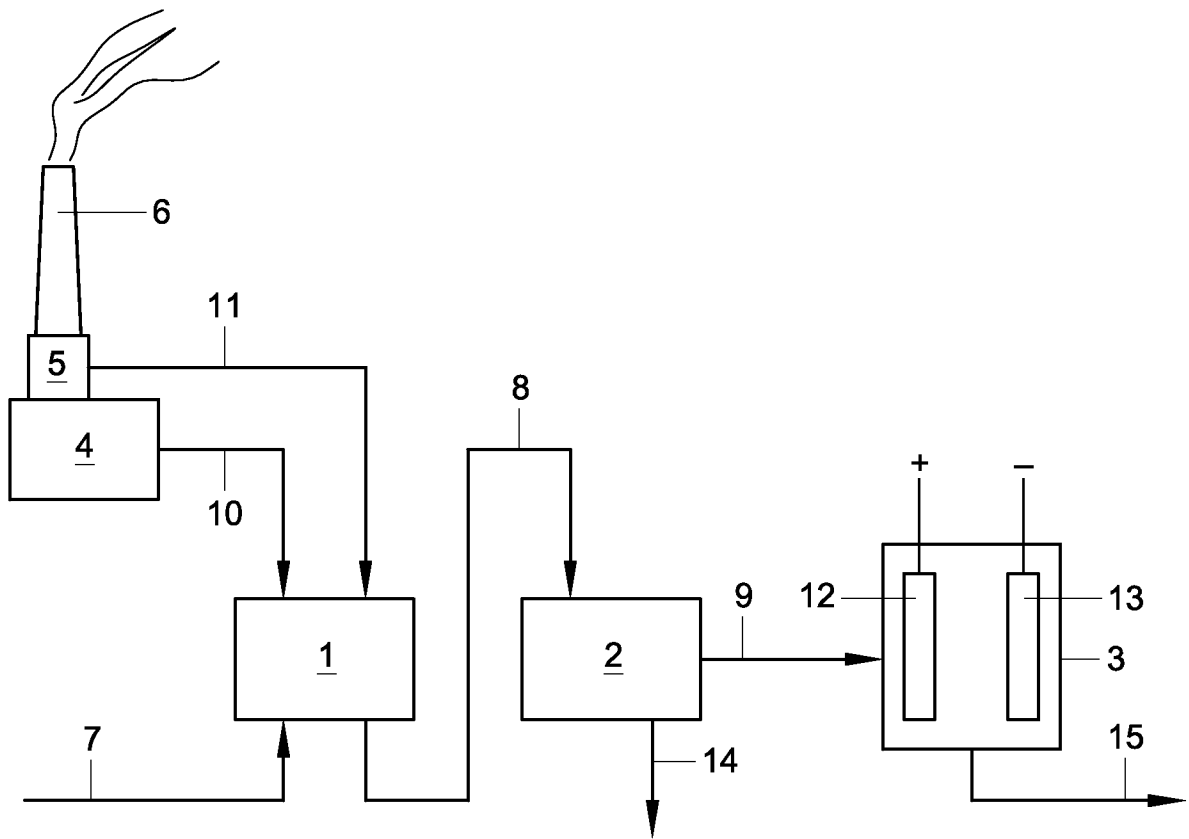


Fig. 2

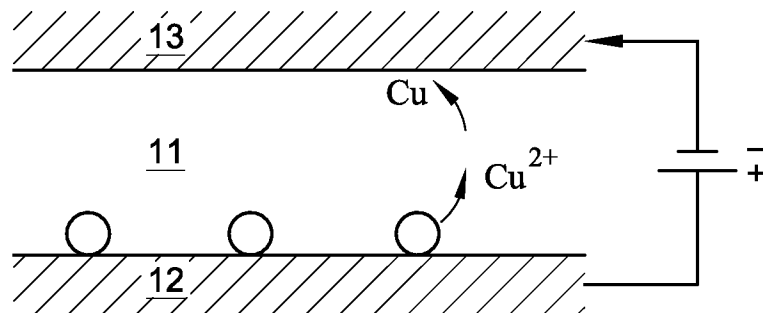


Fig. 1



**INTERNATIONAL SEARCH REPORT**

International application No PCT/NL2010/050657
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**A. CLASSIFICATION OF SUBJECT MATTER**  
 INV. C22B3/02 C22B3/04 C22B3/20 C22B7/00 C22B3/26  
 C25C1/00 C25C7/00 C22B15/00  
 ADD.  
 According to International Patent Classification (IPC) or to both national classification and IPC

**B. FIELDS SEARCHED**  
 Minimum documentation searched (classification system followed by classification symbols)  
 C22B C25C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)  
 EPO-Internal, WPI Data

**C. DOCUMENTS CONSIDERED TO BE RELEVANT**

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	WO 98/58090 A2 (WAART MICHAEL WILHELMUS HENDRI [NL]) 23 December 1998 (1998-12-23) cited in the application claims 1,12; figure 1	1-14
X	GB 1 491 851 A (KRUPP GMBH) 16 November 1977 (1977-11-16) page 2; claim 1; figure 1; examples 1,2	13
Y	DE 33 09 772 A1 (GOEMA DR GOETZELMANN KG PHYSIK [DE]) 20 September 1984 (1984-09-20) page 4, line 32 - page 5, line 1; claim 4; figures 1-3; example 3	1-12,14
Y		1-14
	-/--	

Further documents are listed in the continuation of Box C.

See patent family annex.

\* Special categories of cited documents :

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Date of the actual completion of the international search  11 January 2011	Date of mailing of the international search report  21/01/2011
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## INTERNATIONAL SEARCH REPORT

International application No

PCT/NL2010/050657

C(Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	GB 2 176 207 A (ATOMIC ENERGY AUTHORITY UK) 17 December 1986 (1986-12-17) page 1, line 6 - line 12; claims 1-7; figure 1 page 2, line 8 - line 39	1-14
Y	FR 2 826 878 A1 (GAZ DE FRANCE [FR]) 10 January 2003 (2003-01-10) claims 1-4	1,3,11
Y	EP 0 616 039 A1 (US VANADIUM CORP [US]) 21 September 1994 (1994-09-21) claims 1-9; example III	1,3,11
Y	EP 0 274 059 A2 (RECYTEC SA [CH]) 13 July 1988 (1988-07-13) column 4 - column 6; claims 4-6,8; figures 1,2	4,5,8-11
X	WO 98/14623 A1 (INT CURATOR RESOURCES LTD [CA]; DREISINGER DAVID [CA]; KANE TERRANCE P) 9 April 1998 (1998-04-09) page 12; claims 1-10	1-3,10
Y	page 12; claims 1-10	1-14
Y	EP 0 654 547 A1 (DESARROLLO MINERO LTDA [CL]) 24 May 1995 (1995-05-24) claims 1-7	4,5,9
Y	EP 1 354 968 A2 (STARCK H C GMBH [DE]) 22 October 2003 (2003-10-22) * abstract; figure 1	4,5,14
X	US 3 929 598 A (STERN WILLIAM R ET AL) 30 December 1975 (1975-12-30)	1,13
Y	* abstract; figure 1	1-14
Y	AT 404 258 B (AVR ABFALLVERWERTUNGS UND ROHS [AT]) 27 October 1998 (1998-10-27) page 2 - page 3; figures 1,2	1,2,6-8, 11,13
X	US 5 354 358 A (LITZ JOHN [US] ET AL) 11 October 1994 (1994-10-11) claim 22; figure 1	1,2,6, 8-10
Y	DATABASE WPI Week 200266 Thomson Scientific, London, GB; AN 2002-611880 XP002616117, & JP 2002 166241 A (TAIHEIYO CEMENT CORP) 11 June 2002 (2002-06-11) * abstract	3,6-8

# INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No PCT/NL2010/050657
---

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
WO 9858090	A2	23-12-1998	AT 205551 T 15-09-2001
			AU 7941398 A 04-01-1999
			CA 2294373 A1 23-12-1998
			DE 69801662 D1 18-10-2001
			DE 69801662 T2 20-06-2002
			EP 0991785 A2 12-04-2000
			ES 2164431 T3 16-02-2002
			NL 1006340 C2 21-12-1998
			US 6355155 B1 12-03-2002
<hr/>			
GB 1491851	A	16-11-1977	DE 2513689 A1 07-10-1976
			FR 2305501 A1 22-10-1976
<hr/>			
DE 3309772	A1	20-09-1984	NONE
<hr/>			
GB 2176207	A	17-12-1986	NONE
<hr/>			
FR 2826878	A1	10-01-2003	NONE
<hr/>			
EP 0616039	A1	21-09-1994	AT 167527 T 15-07-1998
			CA 2117155 A1 16-09-1994
			DE 69411056 D1 23-07-1998
			DE 69411056 T2 07-01-1999
			JP 3507540 B2 15-03-2004
			JP 6345434 A 20-12-1994
			US 6306356 B1 23-10-2001
			<hr/>
EP 0274059	A2	13-07-1988	AU 597464 B2 31-05-1990
			AU 1042088 A 30-06-1988
			BG 60505 B1 30-06-1995
			BR 8707567 A 21-02-1989
			CA 1323854 C 02-11-1993
			CS 8709084 A2 14-08-1990
			DD 264716 A5 08-02-1989
			DE 3776638 D1 19-03-1992
			DK 406688 A 01-09-1988
			WO 8804476 A1 16-06-1988
			ES 2030417 T3 01-11-1992
			FI 883698 A 09-08-1988
			GR 3004483 T3 31-03-1993
			HU 202017 B 28-01-1991
			IE 59264 B1 26-01-1994
			IL 84741 A 16-08-1991
			IN 172198 A1 01-05-1993
			JP 63197592 A 16-08-1988
			MX 168761 B 07-06-1993
<hr/>			
EP 0274059	A2		OA 8900 A 31-10-1989
			PL 269370 A1 18-08-1988
			PT 86348 A 17-01-1989
			RO 106045 B1 30-01-1993
			SU 1621818 A3 15-01-1991
			TR 23132 A 10-04-1989
			US 4874486 A 17-10-1989
			YU 224187 A1 28-02-1989
			ZA 8709342 A 26-04-1989
			<hr/>
WO 9814623	A1	09-04-1998	NONE

# INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No

PCT/NL2010/050657

Patent document cited in search report	Publication date	Patent family member(s)	Publication date																																										
EP 0654547	A1	24-05-1995	<table style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 10%;">AU</td> <td style="width: 40%;">7893594 A</td> <td style="width: 50%;">01-06-1995</td> </tr> <tr> <td>BG</td> <td>99199 A</td> <td>29-09-1995</td> </tr> <tr> <td>BR</td> <td>9404672 A</td> <td>11-07-1995</td> </tr> <tr> <td>CA</td> <td>2136052 A1</td> <td>23-05-1995</td> </tr> <tr> <td>CN</td> <td>1109109 A</td> <td>27-09-1995</td> </tr> <tr> <td>CO</td> <td>4440448 A1</td> <td>07-05-1997</td> </tr> <tr> <td>CZ</td> <td>9402816 A3</td> <td>12-07-1995</td> </tr> <tr> <td>FI</td> <td>945460 A</td> <td>23-05-1995</td> </tr> <tr> <td>JP</td> <td>7207484 A</td> <td>08-08-1995</td> </tr> <tr> <td>NO</td> <td>944446 A</td> <td>23-05-1995</td> </tr> <tr> <td>PL</td> <td>305940 A1</td> <td>29-05-1995</td> </tr> <tr> <td>SI</td> <td>9400415 A</td> <td>30-06-1995</td> </tr> <tr> <td>SK</td> <td>141794 A3</td> <td>13-09-1995</td> </tr> <tr> <td>ZA</td> <td>9409210 A</td> <td>04-08-1995</td> </tr> </table>	AU	7893594 A	01-06-1995	BG	99199 A	29-09-1995	BR	9404672 A	11-07-1995	CA	2136052 A1	23-05-1995	CN	1109109 A	27-09-1995	CO	4440448 A1	07-05-1997	CZ	9402816 A3	12-07-1995	FI	945460 A	23-05-1995	JP	7207484 A	08-08-1995	NO	944446 A	23-05-1995	PL	305940 A1	29-05-1995	SI	9400415 A	30-06-1995	SK	141794 A3	13-09-1995	ZA	9409210 A	04-08-1995
AU	7893594 A	01-06-1995																																											
BG	99199 A	29-09-1995																																											
BR	9404672 A	11-07-1995																																											
CA	2136052 A1	23-05-1995																																											
CN	1109109 A	27-09-1995																																											
CO	4440448 A1	07-05-1997																																											
CZ	9402816 A3	12-07-1995																																											
FI	945460 A	23-05-1995																																											
JP	7207484 A	08-08-1995																																											
NO	944446 A	23-05-1995																																											
PL	305940 A1	29-05-1995																																											
SI	9400415 A	30-06-1995																																											
SK	141794 A3	13-09-1995																																											
ZA	9409210 A	04-08-1995																																											
EP 1354968	A2	22-10-2003	<table style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 10%;">AT</td> <td style="width: 40%;">435931 T</td> <td style="width: 50%;">15-07-2009</td> </tr> <tr> <td>CA</td> <td>2425504 A1</td> <td>17-10-2003</td> </tr> <tr> <td>CN</td> <td>1451787 A</td> <td>29-10-2003</td> </tr> <tr> <td>DE</td> <td>10216944 A1</td> <td>06-11-2003</td> </tr> <tr> <td>DK</td> <td>1354968 T3</td> <td>07-09-2009</td> </tr> <tr> <td>ES</td> <td>2327399 T3</td> <td>29-10-2009</td> </tr> <tr> <td>JP</td> <td>2005036245 A</td> <td>10-02-2005</td> </tr> <tr> <td>MX</td> <td>PA03003341 A</td> <td>22-10-2003</td> </tr> <tr> <td>PT</td> <td>1354968 E</td> <td>21-08-2009</td> </tr> <tr> <td>RU</td> <td>2330125 C2</td> <td>27-07-2008</td> </tr> <tr> <td>SI</td> <td>1354968 T1</td> <td>31-10-2009</td> </tr> <tr> <td>US</td> <td>2003196902 A1</td> <td>23-10-2003</td> </tr> <tr> <td>US</td> <td>2007256931 A1</td> <td>08-11-2007</td> </tr> </table>	AT	435931 T	15-07-2009	CA	2425504 A1	17-10-2003	CN	1451787 A	29-10-2003	DE	10216944 A1	06-11-2003	DK	1354968 T3	07-09-2009	ES	2327399 T3	29-10-2009	JP	2005036245 A	10-02-2005	MX	PA03003341 A	22-10-2003	PT	1354968 E	21-08-2009	RU	2330125 C2	27-07-2008	SI	1354968 T1	31-10-2009	US	2003196902 A1	23-10-2003	US	2007256931 A1	08-11-2007			
AT	435931 T	15-07-2009																																											
CA	2425504 A1	17-10-2003																																											
CN	1451787 A	29-10-2003																																											
DE	10216944 A1	06-11-2003																																											
DK	1354968 T3	07-09-2009																																											
ES	2327399 T3	29-10-2009																																											
JP	2005036245 A	10-02-2005																																											
MX	PA03003341 A	22-10-2003																																											
PT	1354968 E	21-08-2009																																											
RU	2330125 C2	27-07-2008																																											
SI	1354968 T1	31-10-2009																																											
US	2003196902 A1	23-10-2003																																											
US	2007256931 A1	08-11-2007																																											
US 3929598	A	30-12-1975	NONE																																										
AT 404258	B	27-10-1998	NONE																																										
US 5354358	A	11-10-1994	NONE																																										
JP 2002166241	A	11-06-2002	JP 3780358 B2 31-05-2006																																										