

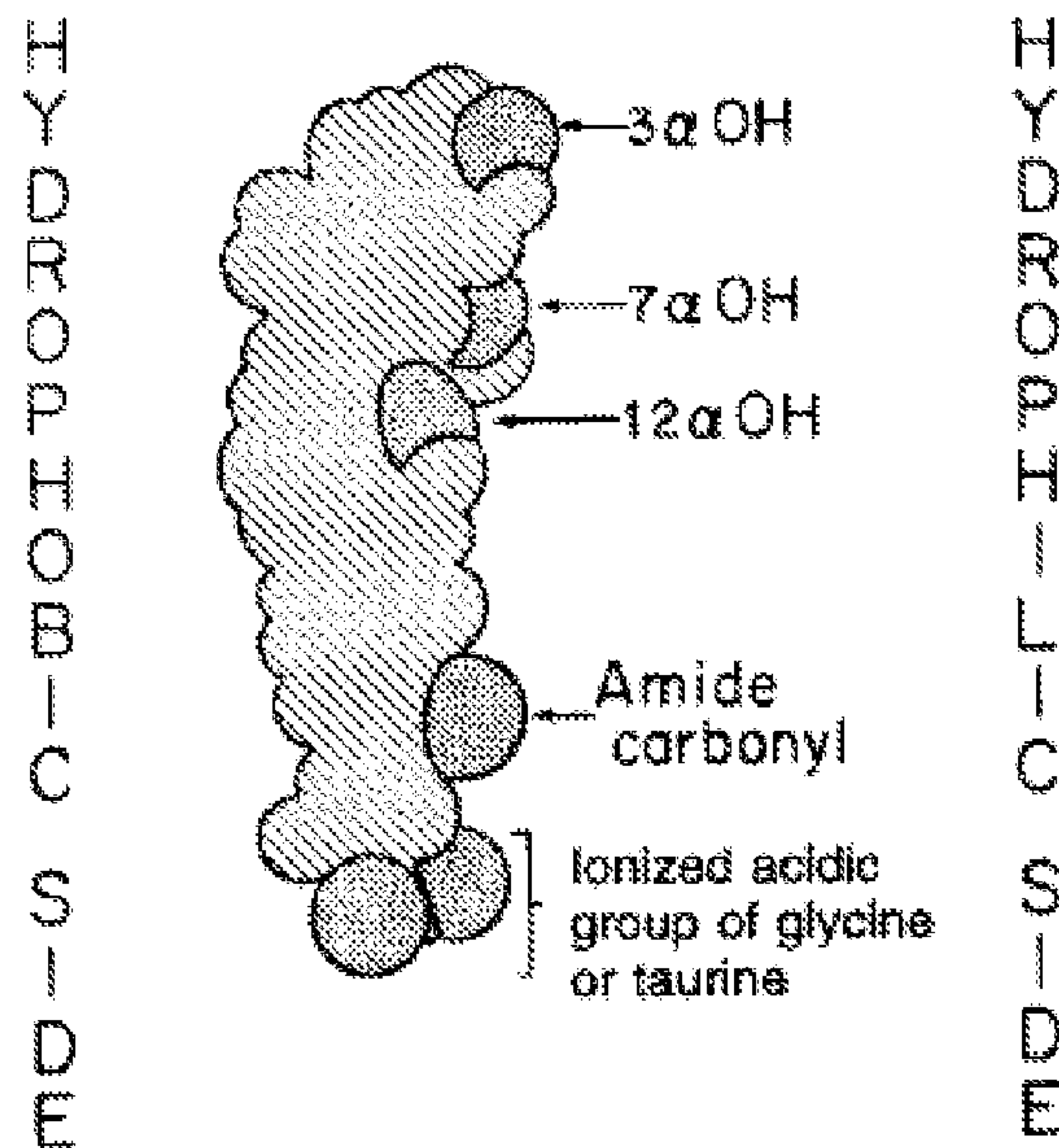


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(54) Titre : COMPOSITIONS D'ACIDES BILIAIRES SYNTHETIQUES ET PROCEDES  
(54) Title: SYNTHETIC BILE ACID COMPOSITIONS AND METHODS

FIGURE 1



(57) Abrégé/Abstract:

Bile acids and related compositions and methods of synthesis and use. More specifically, deoxycholic acid and related compositions, said compositions being free of all moieties of animal origin and free of pyrogenic moieties.

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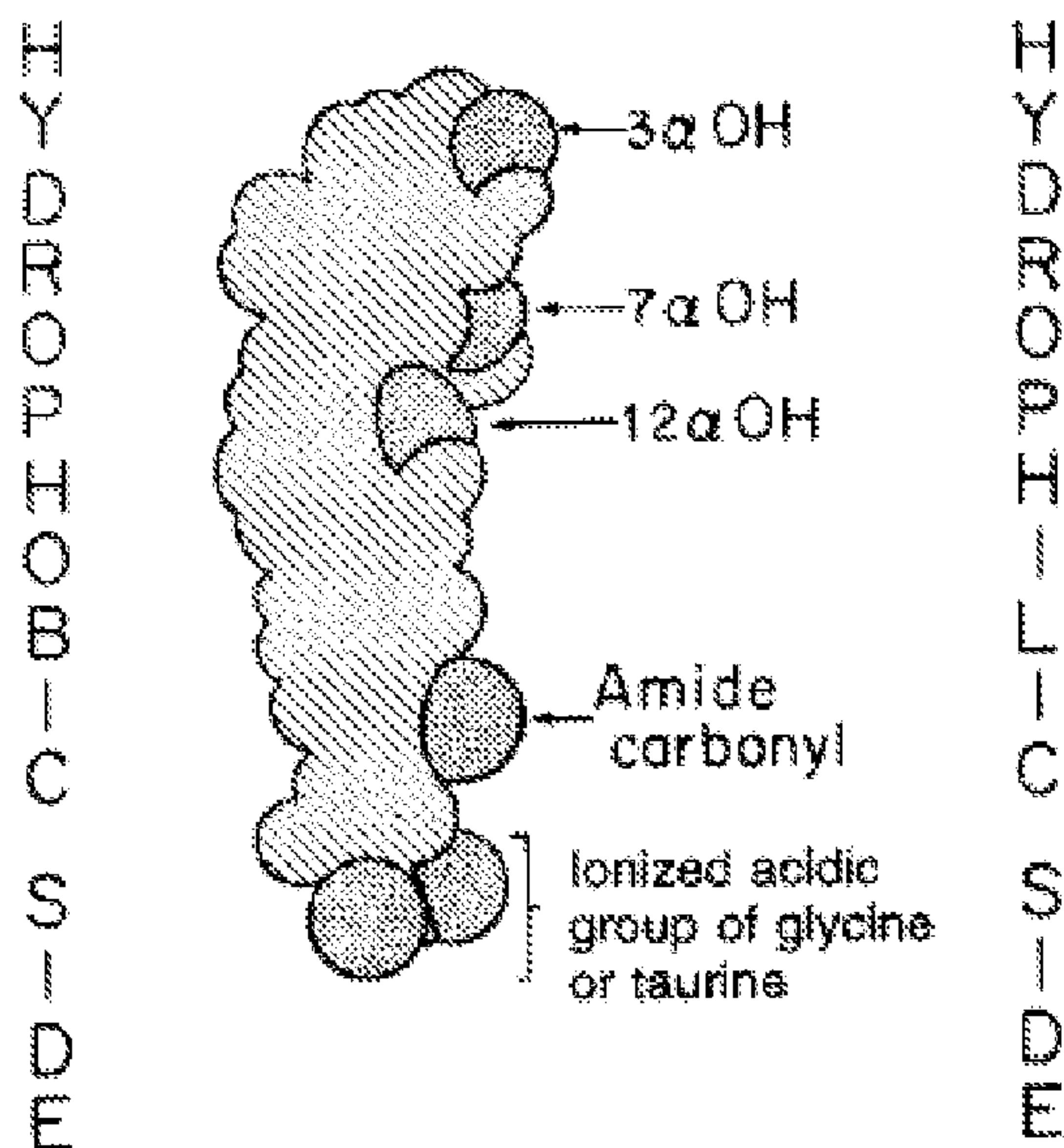
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(54) Title: SYNTHETIC BILE ACID COMPOSITIONS AND METHODS

(57) Abstract: Bile acids and related compositions and  
methods of synthesis and use. More specifically, deoxy-  
cholic acid and related compositions, said compositions  
being free of all moieties of animal origin and free of pyro-  
genic moieties.

FIGURE 1



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## SYNTHETIC BILE ACID COMPOSITIONS AND METHODS

### Field of the Invention

[0001] The present invention relates broadly to bile acids and related compositions and methods. In one aspect, the present invention relates to deoxycholic acid and related compositions, useful intermediates, and methods for synthesis thereof. In another aspect, the present invention relates to use of the present compositions and methods as pharmaceutical compositions as well as methods for the manufacture thereof. Importantly, the bile acids of the present invention are not isolated from mammalian and microbial organisms naturally producing these acids and thus are free of any toxins and contaminants associated with such organisms.

### Background of the Invention

[0002] Cholanology, the study of bile acids, and particularly bile acid chemistry has been of interest for the better part of a century. Although much is known, bile acid chemistry involves a wide variety of chemical entities, many with surprising properties. For a review, *see, e.g.*, Mukhopadhyay, S. and U. Maitra., Current Science 87: 1666-1683 (2004) (“Chemistry and biology of bile acids”), incorporated herein by reference.

[0003] Bile acids are characterized by two connecting units, a rigid steroid nucleus and a short aliphatic side chain (see Figure 1 of the present application). *See*, Hofmann, A. F., et al. For a proposed nomenclature for bile acids, see J. Lipid Res. 33:599-604 (1992). Both the nucleus and the side chain have a large number of possible steric arrangements. The nucleus can be altered by expansion or contraction of individual rings, and the side chain can be shortened or lengthened. In addition, both parts of the bile acid molecule have a large number of possible polar substituents. Ionizing groups may be present on the nucleus or the side chain. Finally, conjugating groups may be present on the nucleus (e.g., sulfate, glucuronate, phosphate) or on the side chain (glycine or taurine or other amino acids, or even sugars). The side chain structure determines the class of the compound (bile acids or bile salts).

[0004] Bile acids are amphiphiles, having both an amphiphilic and amphipathic “face” as shown in Figure 1 (Hofman, A.F., News Physiol. Sci. 14: 24-29 (1999) (“Bile Acids: The Good, the Bad, and the Ugly”, at 25, Figure 1).

[0005] By convention, the hydrophobic surface is called the “ $\beta$ -face” and the hydrophilic surface is called the “ $\alpha$ -face”. The  $\beta$ -face is lipid soluble and the  $\alpha$ -face is relatively polar, in general. There are bile acids, such as those having polar groups (hydroxyl groups, in naturally occurring bile acids) on the hydrophobic face as well as on the hydrophilic face, e.g., ursodeoxycholic acid. The amphipathic nature of the molecule is responsible for its forming mixed micelles with amphipathic but water-insoluble lipids, such as phosphatidylcholine. Bile acids will not solubilize dietary lipids in the form of mixed micelles unless bile acids are above a critical concentration, termed the critical micellization concentration.

[0006] The bile acids found in greatest proportion in humans are chenodeoxycholic acid and deoxycholic acid. Deoxycholic acid is also known as deoxycholate, cholanoic acid, and  $3\alpha,12\alpha$ -dihydroxy- $5\beta$ -cholanate. In the human body deoxycholic acid is used in the emulsification of fats for the absorption in the intestine. In research, deoxycholic acid is used as a mild detergent for the isolation of membrane associated proteins. When substantially pure, deoxycholic acid is a white to off-white crystalline powder form. Deoxycholic acid is one of the four main acids produced by the liver. It is soluble in alcohol and acetic acid. The CAS number for deoxycholic acid is [83-44-3].

[0007] Rapid removal of body fat is an age-old ideal, and many substances have been claimed to accomplish such results, although few have shown results. “Mesotherapy”, or the use of injectables for the removal of fat, is not widely accepted among medical practitioners due to safety and efficacy concerns, although homeopathic and cosmetic claims have been made since the 1950’s. Mesotherapy was originally conceived in Europe as a method of utilizing cutaneous injections containing a mixture of compounds for the treatment of local medical and cosmetic conditions. Although mesotherapy was traditionally employed for pain relief, its cosmetic applications, particularly fat and cellulite removal, have recently received attention in the United States. One such reported treatment for localized fat reduction, which was popularized in Brazil and uses injections of phosphatidylcholine, has been erroneously considered synonymous with mesotherapy. Despite its attraction as a purported “fat-dissolving” injection, the safety and efficacy of these cosmetic treatments remain ambiguous to most patients and physicians. *See*, Rotunda, A.M. and M. Kolodney, *Dermatologic Surgery* 32 :, 465–480



(2006) (“Mesotherapy and Phosphatidylcholine Injections: Historical Clarification and Review”).

**[0008]** WO 2006/133160 (incorporated herein by reference in its entirety including figures) describes methods for lipomodeling, e.g., reduction of a fat depot, by administering a neuropeptide Y receptor antagonist to the site of the fat depot. Kolonin M. G. *et al.*, Nat. Med. June 10(6):625-32 (2004), describes fat selective pro-apoptotic peptides having potent fat cell killing effects. The described pro-apoptotic peptides require access to the vasculature to kill.

**[0009]** Recently published literature reports that deoxycholic acid has fat removing properties when injected into fatty deposits *in vivo*. See, WO 2005/117900 and WO 2005/112942, as well as US2005/0261258; US2005/0267080; US2006/127468; and US20060154906, all incorporated herein by reference in their entirety including figures). Deoxycholate injected into fat tissue has two effects: 1) it kills fat cells via a cytolytic mechanism; and 2) it causes skin tightening. Both of these effects are required to mediate the desired aesthetic corrections (i.e., body contouring). Because deoxycholate injected into fat is rapidly inactivated by exposure to protein and then rapidly returns to the intestinal contents, its effects are spatially contained. As a result of this attenuation effect that confers clinical safety, fat removal therapies typically require 4 – 6 sessions. This localized fat removal without the need for surgery is beneficial not only for therapeutic treatment relating to pathological localized fat deposits (e.g., dyslipidemias incident to medical intervention in the treatment of HIV), but also for cosmetic fat removal without the attendant risk inherent in surgery (e.g., liposuction). See, Rotunda *et al.*, Dermatol. Surgery 30: 1001-1008 (2004) (“Detergent effects of sodium deoxycholate are a major feature of an injectable phosphatidylcholine formulation used for localized fat dissolution”) and Rotunda *et al.*, J. Am. Acad. Dermatol. (2005: 973-978) (“Lipomas treated with subcutaneous deoxycholate injections”), both incorporated herein by reference.

**[0010]** Pharmaceutical grade bile acid preparations are commercially available at relatively low cost. This low cost is due to the fact that the bile acids are obtained from animal carcasses, particularly large animals such as cows and sheep. Importantly, as with all medicaments from animal sources, there is concern that the animal-derived bile acid

products may contain animal pathogens and other harmful agents such as animal or microbial metabolites and toxins, including bacterial toxins such as pyrogens.

**[0011]** Such animal pathogens can include prions, which are thought to be a type of infectious pathogenic protein that may cause prion diseases. Prion diseases are degenerative disorders of the nervous system. One such disease, “Mad cow” disease (thought to be a variant of Creutzfeldt-Jakob disease (CJD)), is thought to be caused by a prion present in edible beef from diseased cows. Most cases are sporadic with unknown mode of transmission; some cases are inherited; and a small number have been transmitted by medical procedures. The spread of human prion diseases through consumption of infected material has been implicated historically in kuru and recently in variant CJD. Other animal prion diseases (scrapie of sheep, transmissible mink encephalopathy, chronic wasting disease of cervids, and bovine spongiform encephalopathy) all seem to be laterally transmitted by contact with infected animals or by consumption of infected feed. Risk assessment and predictions of future events pertaining to prion diseases are difficult to ascertain because of the different modes of transmission, the unpredictable species barriers, the variable distribution of infectivity in tissues, and strain variations found in some diseases.

**[0012]** In general, animal products may be exposed to microbial organisms which produce pyrogens (fever-causing substances). Bacterial contaminants of food and/or pharmaceutical products are also a serious issue as evidenced by contamination of food stuffs by enterohemorrhagic *E. coli*. Products such as meats derived from cows as well as produce such as apples, spinach, and the like, have been implicated in such contamination. In such cases, it is the toxin produced by the bacteria (rather than the bacteria itself) that produces adverse effects in humans. Such adverse effects include severe diarrhea, kidney failure and in the extreme situations, death. Bacterial endotoxins, a type of pyrogen, must be substantially excluded from all pharmaceutical compositions.

**[0013]** Animal products are generally purified by a process of elimination, i.e., rather than selecting the end-product from a mix, the end product is the material remaining after exclusion of impurities. And, in addition to the potential animal moieties such as pathogens, another artifact of purification from animal sources is that the end-product is a mixture of one or more bile acids. For example, commercial preparations of deoxycholic acid contain some chenodeoxycholic acid, as well as cholic acid, which is a precursor to



both deoxycholic acid and chenodeoxycholic acid in mammalian bile acid synthesis. Because the exact proportion of deoxy/cheno/cholic is not preselected, this may result in lot-to-lot variation when contemplating manufacturing large amounts of bile acids. Such lot-to-lot variation can be problematic and may engender additional steps in garnering regulatory approvals or quality control, particularly in efforts to produce a pharmaceutical composition. Clearly, producers would desire lot-to-lot predictability in manufacturing bile acid pharmaceutical compositions.

**[0014]** Currently, the concerns regarding animal-derived products containing animal pathogens and other harmful agents has been addressed by sourcing from isolated and inspected animals. For example, deoxycholic acid from animals in New Zealand are a source of bile acids for human use under US regulatory regimes, as long as the animals continue to remain isolated and otherwise free of observable pathogens.

**[0015]** Implicitly, by the need for such governmentally controlled regulatory regime is the recognition of an intrinsic risk of transmission of animal pathogens when animal-derived medicaments are injected. Where non-animal medicament alternatives become available, the governmental regulatory regime is no longer needed. An example of such alternative (non-animal medicament replacing animal-derived medicament) and associated advantages is insulin for human use. The manufacture of beef insulin in the United States was discontinued in 1998, and pork insulin for human use was discontinued in January of 2006. Although animal insulin can be obtained from herds not known to have had exposure to BSE-causing or other pathogenic agents, the manufacturing facilities or processes can expose the animal ingredients to animals which have had exposure to the pathogens. The risk of transmission of pathogenic agents to humans can be eliminated with the use of insulin that is manufactured recombinantly or synthetically. For consumers, the insulin situation is instructive: where synthetic material is freely available, the risk of transmission of animal pathogens is in theory eliminated. For producers, the ability to produce a pure chemical entity that is substantially free of material of animal pathogens is advantageous for safety, quality, and regulatory purposes. Further, a synthetic process typically provides for a more reproducible product than that derived from biological sources.

**[0016]** Presently, because of the relative abundance of animal carcass-derived bile acids, the industry has not taken steps to either fully chemically synthesize bile acids, or



prepare bile acids using phytosterol or microbial starting materials. And although bile acid derivatives have been synthesized, this work again primarily involved animal-derived bile acids as starting materials for steroid chemistry, due to the low cost and ready availability of animal materials. Despite historically active efforts in phytosterol research, there are no readily-commercially available phytosterol-derived bile acid pharmaceutical grade compositions. *See, e.g.,* Mukhopadhyay, S. and U. Maitra., *Current Science* 87: 1666-1683, 1670 (2004) (Noting that the total synthesis of any bile acid had not been performed subject to a 1981 reference, Kametani et al. *J. Am. Chem. Soc.* 103: 2890 (1981) (“First Total Synthesis of (+)-Chenodeoxycholic Acid”). Microbial, such as bacterially-produced bile acids, have been used *in situ* as bacterial products, e.g., for marine oil spill clean-up. *See, Maneerat et al., Appl. Microbiol. Biotechnol.* 76: 679-683 (2004) (“Bile acids are new products of a marine bacterium, *Myroides* sp. Strain SM1”).

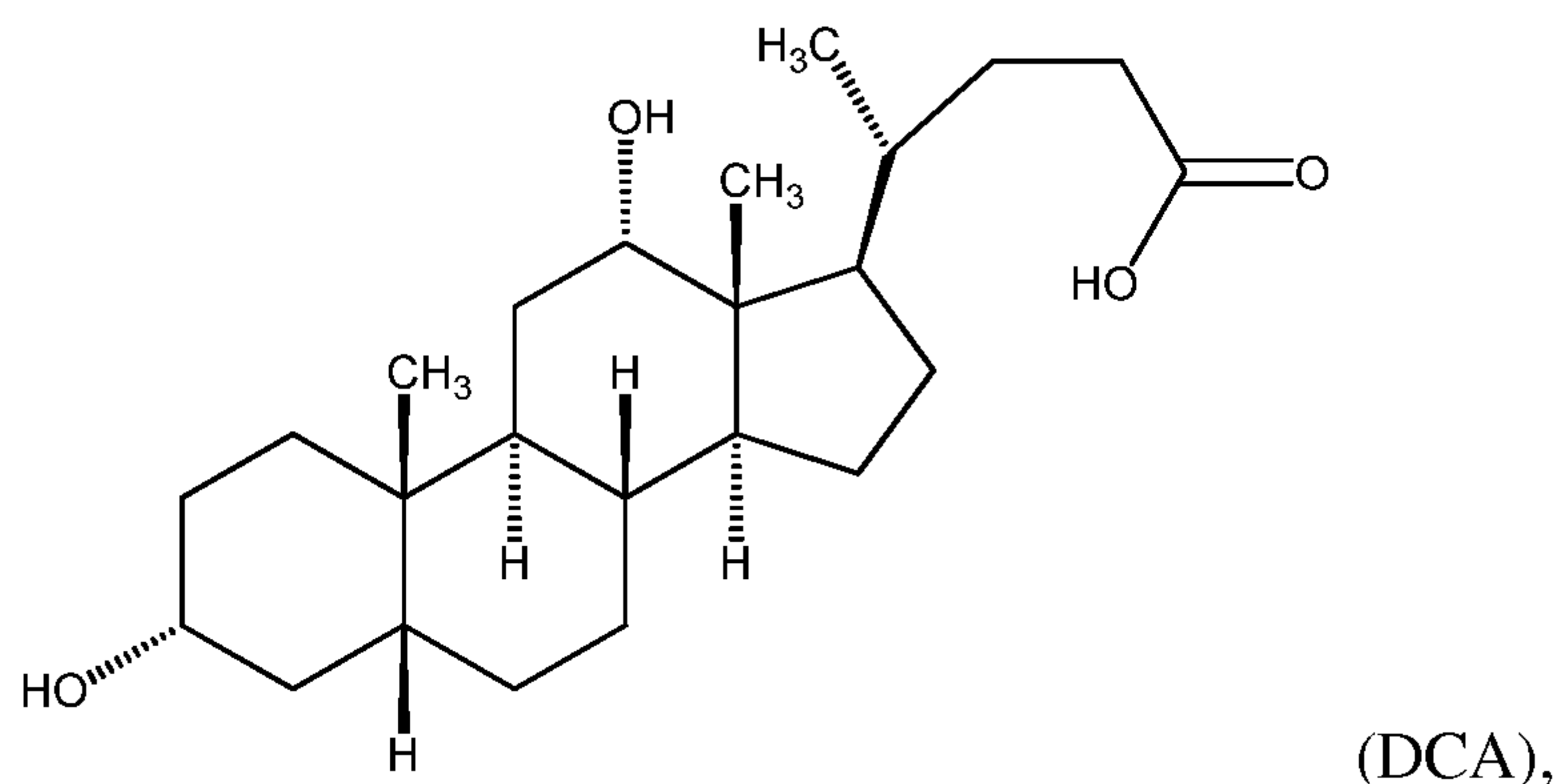
**[0017]** In order to realize the full potential of deoxycholic acid for the removal of fat, it is imperative that the concerns over the use of animal derived products be further addressed. Clearly, there is a need for suitable quantities of efficacious bile acids and related compositions, such as the deoxycholic acids, that are known from the outset to be free from moieties of animal origin (or pathogenic moieties capable of acting in an animal, particularly a mammal, and for human use, having a deleterious effect on a human), and other harmful agents such as animal or microbial metabolites, toxins, including bacterial toxins, such as pyrogens, for use as medicaments in humans. The present invention addresses this concern by providing synthetically prepared bile acid compositions free of the potential risk of animal pathogens and other harmful agents. The disclosed bile acid compositions can be used in adipolytic therapy and will serve to further advance research and developmental efforts in the area of localized fat removal.

### Summary of the Invention

**[0018]** Adequate quantities of suitable bile acid as a defined pharmaceutical composition is herein provided, as well as methods for synthesis thereof. Bile acid compositions and methods so provided are not isolated from mammalian or microbial organisms that naturally produce the bile acids. In one aspect, particular deoxycholic acid pharmaceutical compositions which are free of all moieties of animal origin and of mammalian and/or bacterial pyrogens, and related methods for production and use are provided. In another aspect, adequate quantities of suitable deoxycholic acids as defined

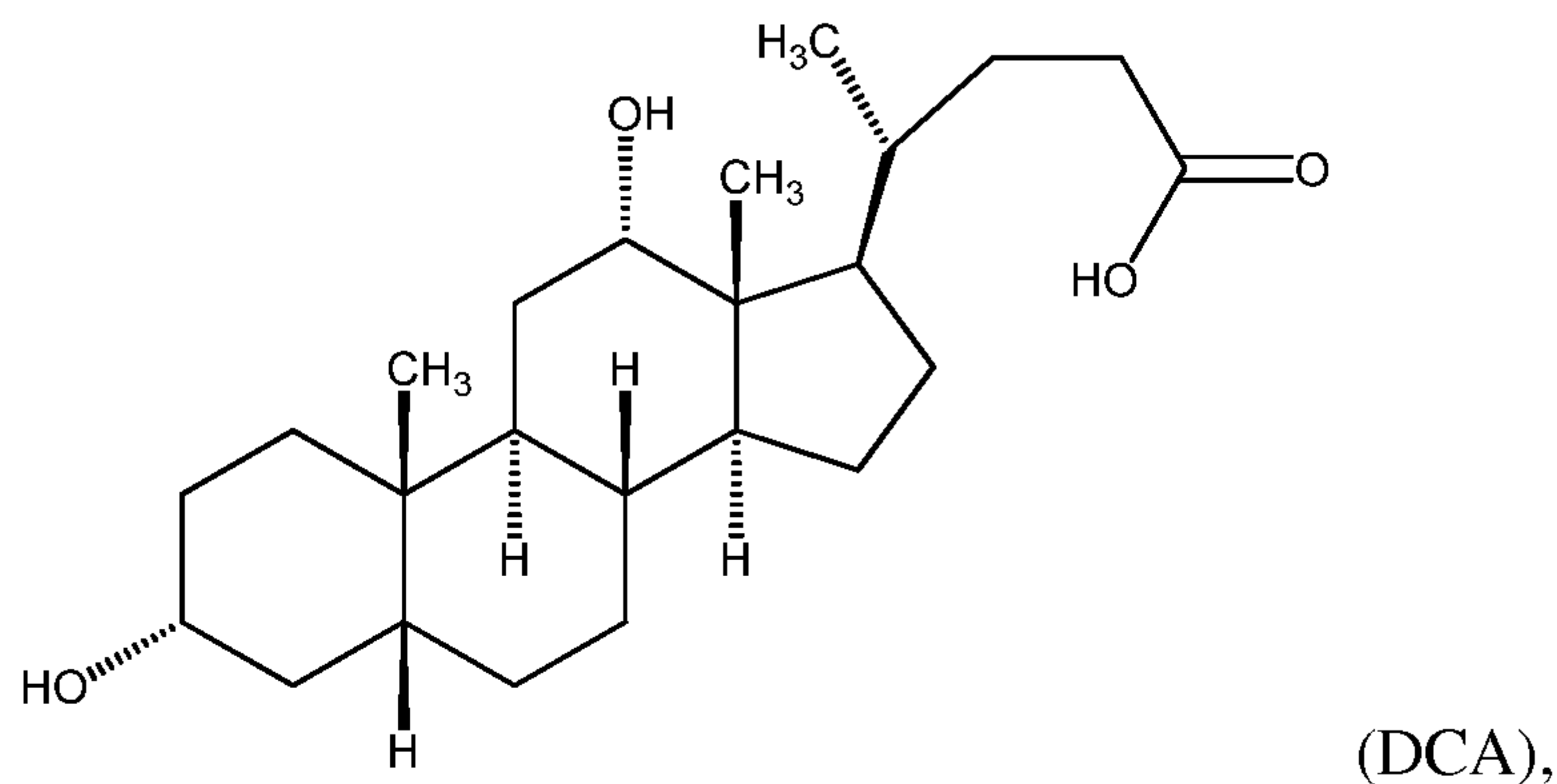
pharmaceutical compositions are provided which can be used as an injectable pharmaceutical composition for localized fat removal, along with related compositions, methods for manufacture and methods of use. The defined deoxycholate injectates of the present invention may be combined with a molecule that causes fat to die by an orthogonal mechanism, e.g., NPY antagonists and/or fat selective pro-apoptotic peptides, to provide agents to be used to create a more potent means to mediate body contouring in fewer therapeutic sessions.

**[0019]** In one aspect, this invention provides a compound that is a synthetic deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



wherein said DCA has a  $^{14}\text{C}$  content of less than 1 ppt.

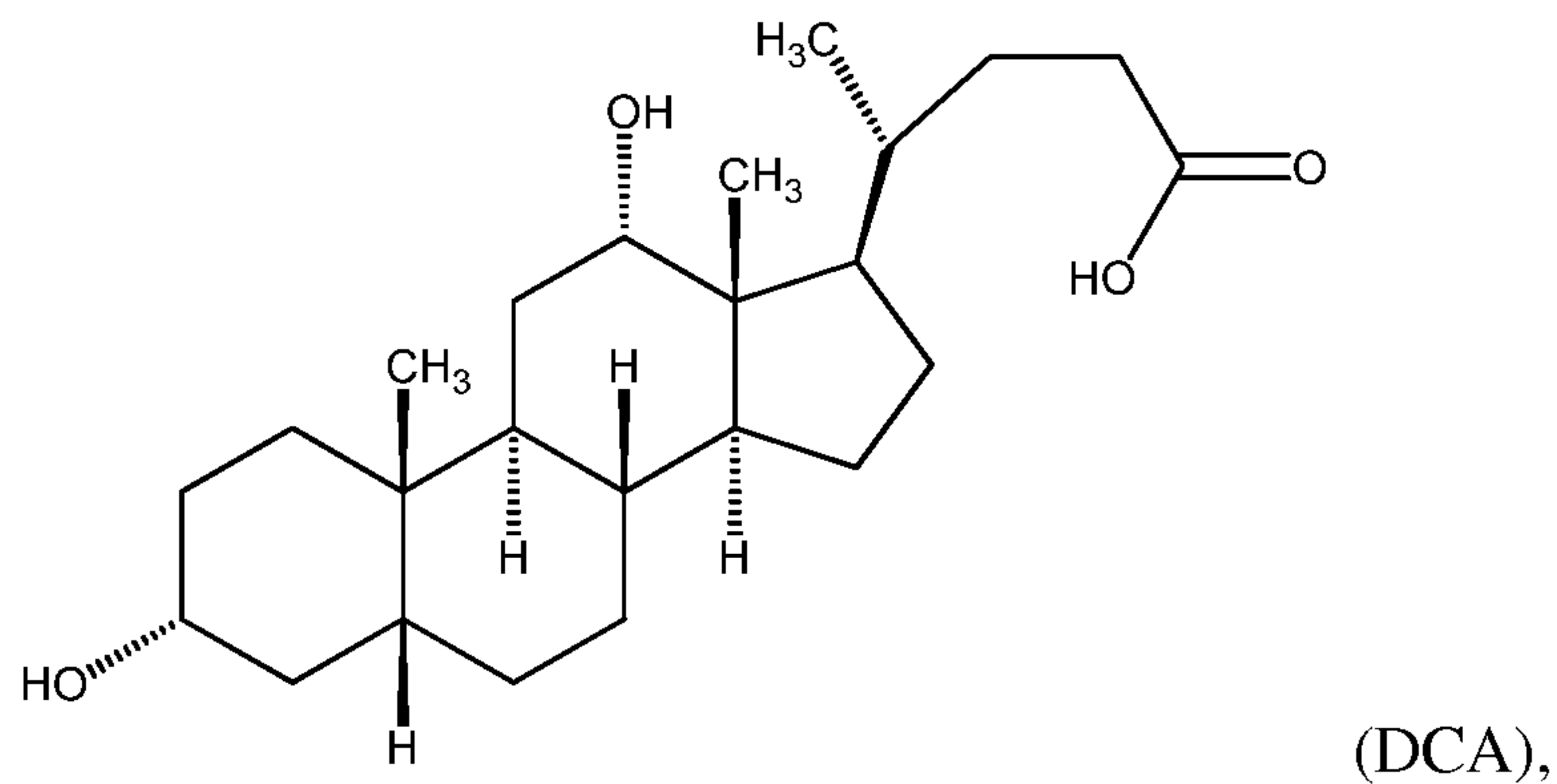
**[0020]** In another aspect, this invention provides a synthetic deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



wherein said DCA has a  $^{14}\text{C}$  content of less than 0.9 ppt.

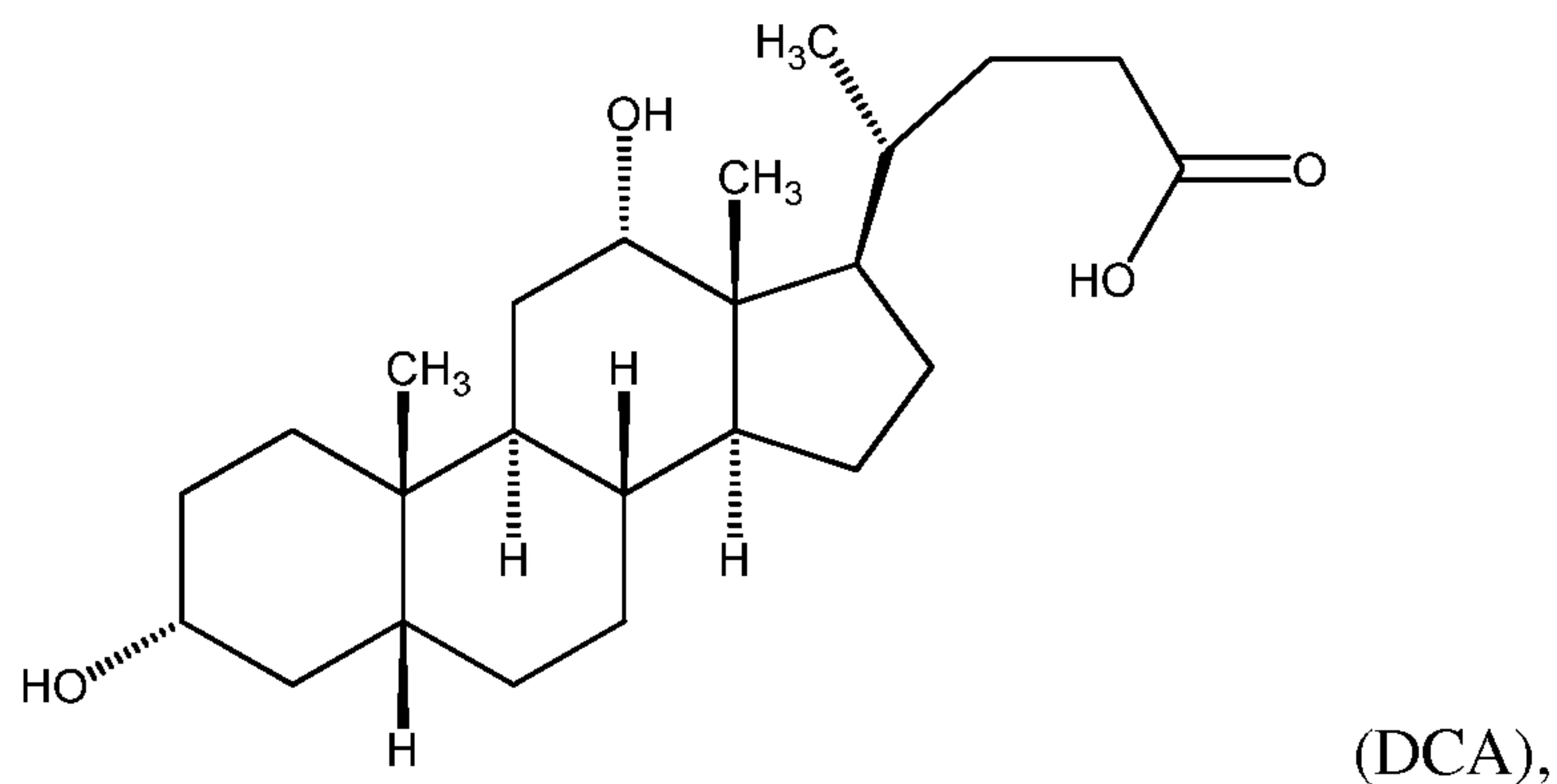


[0021] In another aspect, this invention provides a synthetic deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



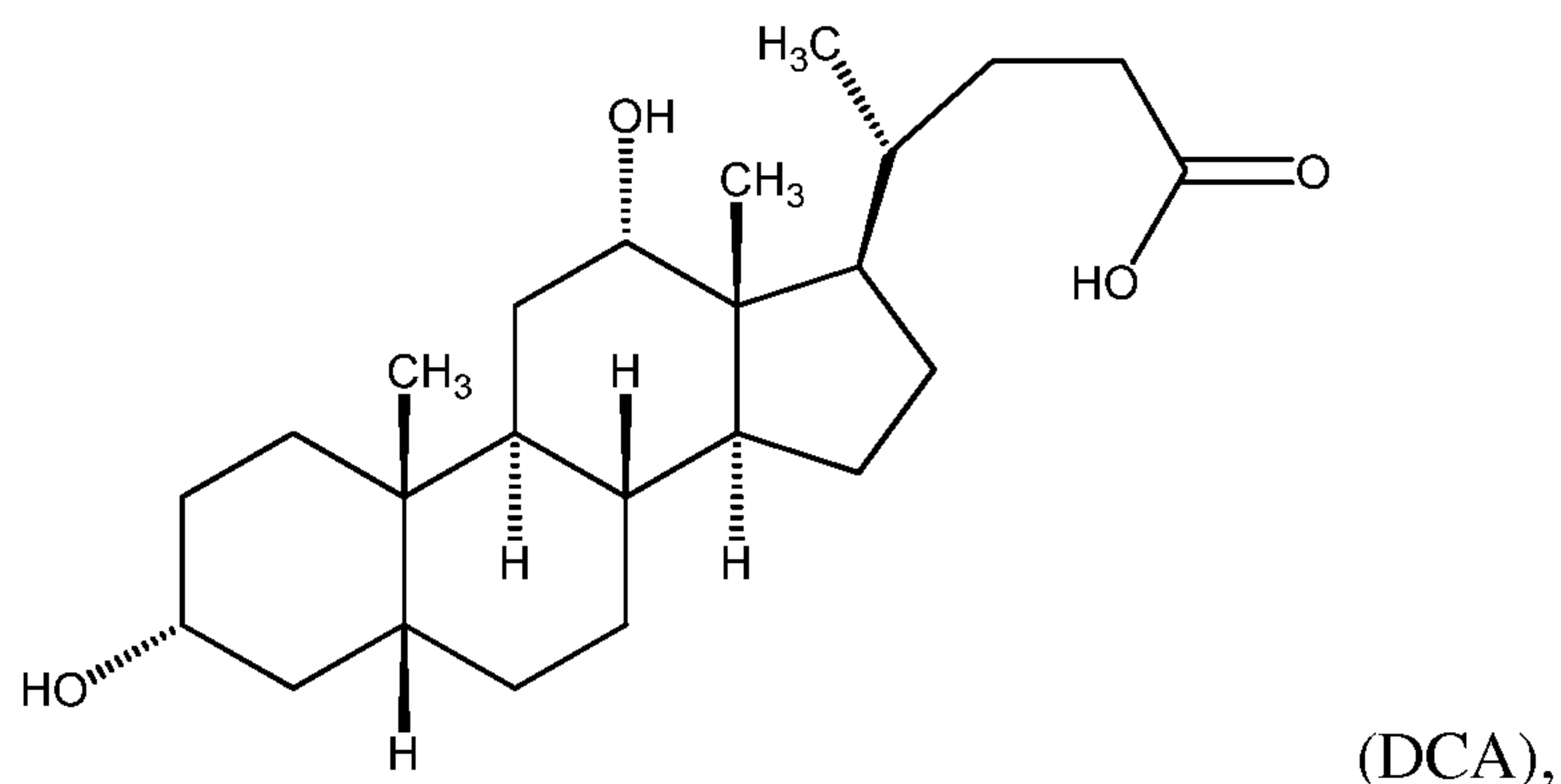
wherein said DCA has a  $^{14}\text{C}$  content of less than 0.86 ppt.

[0022] In another aspect, this invention provides a compound that is a synthetic deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



wherein said DCA has a  $^{14}\text{C}$  content of greater than 1 ppt.

[0023] In still another aspect, this invention provides a compound that is a synthetic deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



wherein said DCA has a  $^{14}\text{C}$  content of 1 ppt.

[0024] In some embodiments, the invention provides the pharmaceutically acceptable salt of the DCA this invention. In some embodiments, the salt is a sodium salt, which is also referred to as sodium deoxycholate.

[0025] In still another aspect, this invention provides a pharmaceutical or cosmetic composition comprising the DCA or a pharmaceutically acceptable salt thereof of this invention and a pharmaceutically acceptable excipient. In some embodiments, the composition is for non-surgical reduction or removal of localized fat deposit in a patient in need thereof.

[0026] In still another aspect, this invention provides a method for determining whether a sample of deoxycholic acid was prepared synthetically, which method comprises:

- a) assaying a sample of deoxycholic acid or a salt thereof for its fossil carbon or  $^{14}\text{C}$  content;
- b) comparing the fossil carbon or  $^{14}\text{C}$  content of the sample against the fossil carbon or  $^{14}\text{C}$  content of naturally occurring deoxycholic acid to determine whether the sample was prepared synthetically.

### **Brief Description of the Drawings**

[0027] Figure 1 illustrates the amphiphilic and amphipathic “faces” of bile acid (Hofman, A.F., News Physiol. Sci. 14: 24-29 (1999) (“Bile Acids: The Good, the Bad, and the Ugly”, at 25, Figure 1).

[0028] Figure 2 is a drawing representing the structure of bile acids, including the numbering system for the carbons of the bile acid skeleton.

### **Incorporation by Reference**

[0029] All publications and patent applications mentioned in this specification are herein incorporated by reference to the same extent as if each individual publication or patent application was specifically and individually indicated to be incorporated by reference.

### **Detailed Description of the Invention**

[0030] Throughout this disclosure, various publications, patents and published patent specifications are referenced by an identifying citation. The disclosures of these



publications, patents and published patent specifications are hereby incorporated by reference into the present disclosure to more fully describe the state of the art to which this invention pertains.

[0031] As used herein, certain terms have the following defined meanings.

[0032] The term “pathogen” refers to a specific causative agent of a disease.

[0033] The term “animal origin” refers to originating from any of a kingdom (Animalia) of living things including many-celled organisms and single celled organisms.

[0034] The term “mammalian origin” refers to originating from any mammalian organism. The term “mammalian organism” refers to a class (Mammalia) of warm-blooded higher vertebrates (as placentals, marsupials, or monotremes) that nourish their young with milk secreted by mammary glands, have the skin usually more or less covered with hair, and include humans.

[0035] The term “microbial origin” refers to originating from any microbial organism. The term “microbial organism” refers to a domain (Bacteria) of prokaryotic round, spiral, or rod-shaped single-celled microorganisms that may lack cell walls or are gram-positive or gram-negative if they have cell walls, that are often aggregated into colonies or motile by means of flagella, that typically live in soil, water, organic matter, or the bodies of plants and animals, that are usually autotrophic, saprophytic, or parasitic in nutrition, and that are noted for their biochemical effects and pathogenicity.

[0036] The term “lower alkyl” refers to monovalent saturated aliphatic hydrocarbyl groups having from 1 to 10 carbon atoms and preferably 1 to 6 carbon atoms. This term includes, by way of example, linear and branched hydrocarbyl groups such as methyl ( $\text{CH}_3-$ ), ethyl ( $\text{CH}_3\text{CH}_2-$ ), *n*-propyl ( $\text{CH}_3\text{CH}_2\text{CH}_2-$ ), isopropyl ( $(\text{CH}_3)_2\text{CH}-$ ), *n*-butyl ( $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2-$ ), isobutyl ( $(\text{CH}_3)_2\text{CHCH}_2-$ ), *sec*-butyl ( $(\text{CH}_3)(\text{CH}_3\text{CH}_2)\text{CH}-$ ), *t*-butyl ( $(\text{CH}_3)_3\text{C}-$ ), *n*-pentyl ( $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2-$ ), and neopentyl ( $(\text{CH}_3)_3\text{CCH}_2-$ ).

[0037] “Aryl” or “Ar” refers to a monovalent aromatic carbocyclic group of from 6 to 14 carbon atoms having a single ring (*e.g.*, phenyl or naphthyl).

[0038] The term “ethanedithiol or dithiane precursor” refers to a reagent that, with reaction with a carbonyl group, will form an ethane dithiol or dithiane group.

[0039] The term “oxidizing agent” refers to a reagent which can accept electrons in an oxidation-reduction reaction. In this way, halogen or oxygen can be added to a molecule or hydrogen can be removed from a molecule.

[0040] The term “desulfurization reagent” refers to a reagent which can react with a sulfide. In one aspect, a desulfurization reagent can react with a sulfide containing molecule to remove the sulfide group from the molecule.

[0041] The term “reducing agent” refers to a reagent which can donate electrons in an oxidation-reduction reaction. In this way, halogen or oxygen can be removed from a molecule or hydrogen can be added to a molecule.

[0042] The term “electrophilic acetyl group” refers to an acetyl group as an electrophile, a group which is attracted to electrons and tends to accept electrons.

[0043] The term “acetylating reagent” refers to a reagent in which can add an acetyl group to a molecule.

[0044] The term “acid” refers to a proton donor.

[0045] The term “hydrogenation reagent” refers to a reagent that can donate hydrogen to a molecule.

[0046] The term “dehydration reagent” refers to a reagent that can react with water. In one aspect, a dehydration reagent can react with water that is removed from a molecule.

[0047] In various aspects described herein, the present invention provides compositions (and useful intermediates) for pharmaceutical use, methods of synthesis thereof, and methods of use of the present pharmaceutical compositions.

[0048] Importantly, the present bile acid compositions are free of risks inherent in material obtained from animal starting materials, and therefore do not require the detailed inspections and regulations of animal-derived materials. In one aspect, this invention is thus directed to bile acid pharmaceutical compositions free of material of animal origin, such as mammalian pathogens, as well as being substantially free of toxins of bacterial origin, such as pyrogens. The present bile acid pharmaceutical compositions are optionally in salt form, and, further optionally contain a pharmaceutically acceptable diluent, excipient or carrier. Cations for salt preparation may be selected from the group consisting of sodium ( $\text{Na}^+$ ), potassium ( $\text{K}^+$ ), lithium ( $\text{Li}^+$ ), magnesium ( $\text{Mg}^{2+}$ ), calcium



(Ca<sup>2+</sup>), barium (Ba<sup>2+</sup>), strontium (Sr<sup>2+</sup>), ammonium (NH<sub>4</sub><sup>+</sup>), primary, secondary and tertiary alkyl and aryl ammonium salts (NR<sup>10</sup>H<sub>3</sub><sup>+</sup>, N(R<sup>10</sup>)<sub>2</sub>H<sub>2</sub><sup>+</sup> and N(R<sup>10</sup>)<sub>3</sub>H<sup>+</sup>, respectively, wherein R<sup>10</sup> is independently alkyl or aryl). Also functional group derivatives of the secondary hydroxyl groups such as esters and ethers. Salts may also be prepared from an alkali metal or an alkaline earth metal. An alkali metal may be selected from among sodium (Na<sup>+</sup>), potassium (K<sup>+</sup>), and lithium (Li<sup>+</sup>). An alkaline earth metal may be selected from the group consisting of magnesium (Mg<sup>2+</sup>), calcium (Ca<sup>2+</sup>), barium (Ba<sup>2+</sup>), and strontium (Sr<sup>2+</sup>). Preferably for use as a pharmaceutical composition for localized removal of fat, the bile salt is sodium deoxycholate.

**[0049]** Prodrugs of the compounds of the embodiments are also contemplated. A prodrug is an active or inactive compound that is modified chemically through *in vivo* physiological action, such as hydrolysis, metabolism and the like, into a compound of the embodiments following administration of the prodrug to a patient. For example, one may prepare an ester of the present deoxycholic acid or derivatives thereof, so that the release of the deoxycholic acid or derivatives thereof is triggered by the disruption of the cell membrane, and release of esterase. One may also use derivatives of the carboxyl group such as esters, amides and peptides. Derivatives of the hydroxyl groups, such as esters and ethers, may also be used. With the release of esterase, the ester protecting group is cleaved so that the deoxycholic acid active form or derivatives thereof is present at the desired location *in situ*. For a general discussion of prodrugs involving esters see Svensson and Tunek Drug Metabolism Reviews 165 (1988) and Bundgaard Design of Prodrugs, Elsevier (1985), herein entirely incorporated by reference.

**[0050]** In general, the compounds of preferred embodiments will be administered in a therapeutically effective amount by any of the accepted modes of administration for agents that serve similar utilities. The actual amount of the compound of preferred embodiments, i.e., the active ingredient, will depend upon numerous factors such as the severity of the disease to be treated, the age and relative health of the subject, the potency of the compound used, the route and form of administration, and other factors. The drug can be administered more than once a day, preferably once or twice a day. All of these factors are within the skill of the attending clinician.

**[0051]** The compositions can be comprised of a disclosed compound in combination with at least one pharmaceutically acceptable excipient. Acceptable excipients are

non-toxic, aid administration, and do not adversely affect the therapeutic benefit of the disclosed compound. Such excipient may be any solid, liquid, semi-solid or, in the case of an aerosol composition, gaseous excipient that is generally available to one of skill in the art.

**[0052]** Solid pharmaceutical excipients include starch, cellulose, talc, glucose, lactose, sucrose, gelatin, malt, rice, flour, chalk, silica gel, magnesium stearate, sodium stearate, glycerol monostearate, sodium chloride, dried skim milk and the like. Liquid and semisolid excipients may be selected from glycerol, propylene glycol, water, ethanol and various oils, including those of petroleum, vegetable or synthetic origin, e.g., peanut oil, soybean oil, mineral oil, sesame oil, etc. Preferred liquid carriers, particularly for injectable solutions, include water, saline, aqueous dextrose, and glycols.

**[0053]** The amount of the compound in a formulation can vary within the full range employed by those skilled in the art. The present compositions in various aspects as described herein may be prepared wherein the deoxycholic acid moiety is in the range of about 0.5%-10% on a weight per aqueous volume basis, or, on a w/w basis assuming the density of water (i.e., a 1:1 correspondence between weight and volume). In another aspect, the present embodiments relate to the presently described pharmaceutical compositions in concentrations up to saturation point of the diluent. One may select the degree of thixotropic viscosity based on conditions such as concentration and pH. *See, e.g.,* Mukhopadhyay, S. and U. Maitra, Current Science 87: 1666-1683 (2004) at 1680.

**[0054]** Of particular note is the potential for local irritation upon injection of a bile acid composition of the present embodiments, and thus it may be desirable to administer, simultaneously or in seriatim, a local anesthetic. For example, lidocaine is frequently used in humans, and may be administered either as a co-formulation (in the same container and injected at the same time) or co-injection (injected from a different container). Anesthetics such as lidocaine may be administered via topical preparation, such as a patch or ointment.

**[0055]** For deeper tissue, anesthetics may be more deeply injected into the subject tissue or administered systemically (e.g., general anesthesia, epidural or other known methods).

**[0056]** The composition of the invention can comprise the deoxycholic acid or salt thereof at a concentration of about 0.001 to 10, 0.01 to 5, or 0.1 to 2% w/w, w/v, or v/v.



Preferably, the deoxycholic acid or salt in the above solution can be at a concentration of about 0.1-5 % w/w or more preferably about 0.5% or 1% w/w. In some embodiments, the fat dissolving solution comprises up to 100, 50, 20, 10, 5, 2, 1, 0.5, 0.2, 0.05, 0.02, or 0.01 grams of the one or more detergents, bile acids and/or bile salts, deoxycholic acid or salts thereof or sodium deoxycholate.

**[0057]** In preferred embodiments, the solutions herein include no lipids, phospholipids, or phosphatidylcholine. In some embodiments, the solutions herein include up to 5% w/w, w/v, or v/v lipids, phospholipids, or phosphatidylcholine. In some embodiments, the lipids, phospholipids, or phosphatidylcholine is in an amount that is less than the deoxycholic acid or the salt thereof in mass.

**[0058]** In some embodiments, the above solution can further comprise a second therapeutic agent selected from the group consisting of: anti-microbial agents, vasoconstrictors, anti-thrombotic agents, anti-coagulation agents, suds-depressants, anti-inflammatory agents, analgesics, dispersion agents, anti-dispersion agents, penetration enhancers, steroids, tranquilizers, muscle relaxants, and anti-diarrhea agents. In some embodiments, a solution is in a container that contains up to 500 mL of solution. Such container can be a syringe or syringe-loadable container.

**[0059]** In some embodiments, compositions and methods further comprise a molecule known to cause fat to die by an orthogonal mechanism. Such molecules include neuropeptide Y (NPY) antagonists including, but not limited to, NPY receptor antagonists, such as BIBP-3226 (Amgen), BIBO-3304 (Boehringer Ingelheim), BMS-192548 and AR-H040922 (Bristol-Myers Squibb), LY-357897 (Eli Lilly), 1229U91 and GW438014S (GlaxoSmithKline), JNJ-5207787 (Johnson & Johnson), Lu-AA-44608 (Lundbeck), MK-0557 (Merck NPY), NGD-95-1 (Neurgogen), NLX-E201 (Neurologix), CGP-71683 (Novartis), PD-160170 (Pfizer), SR-120819A, BIIE0246, and S.A.0204 (Sanofi Aventis), S-2367 (Shiongli), dihydropyridine and dihydropyridine derivatives that are NPY receptor antagonists, bicyclic compounds that are NPY receptor antagonists, carbazole NPY receptor antagonists, and tricyclic compounds that are NPY receptor antagonists. *See, e.g.*, WO 2006/133160 and U.S. 6,313,128 (incorporated herein by reference in its entirety including figures). Also contemplated are fat selective pro-apoptotic peptides such as the CKGGRAKDC peptide that homes to white fat vasculature. *See*, Kolonin M.G. *et al.*, Nat. Med. June 10(6):625-32 (2004).

**[0060]** In some embodiments, the administering step involves delivering the compositions herein via a dermal patch, a pump, or subdermal depot. In some embodiments, the administering step involves delivering the compositions herein topically or subcutaneously. In specific embodiments, the administration step involves administering locally (e.g., subcutaneously or subdermally) to a region under eye, under chin, under arm, buttock, calf, back, thigh, or stomach of said subject. The administration can be made by a subcutaneous or transdermal injection.

**[0061]** In one aspect, the present invention relates to methods for reducing a subcutaneous fat deposit in a subject. Such methods comprise the step of administering locally to a subcutaneous fat deposit in the subject a composition comprising: (i) a fat-dissolving effective amount of one or more pharmacologically active detergents, or bile acid(s) and/or bile salt(s), or deoxycholic acid or a salt thereof, or sodium deoxycholate; (ii) a pharmaceutical, veterinary, or cosmetic excipient; and (iii) optionally a lipid, wherein the ratio of the lipid and bile acid or bile salt is up to 1% w/w and wherein the composition does not include lipase or colipase. In some embodiments, the fat deposit is associated with a condition selected from the group consisting of obesity, fat redistribution syndrome, eyelid fat herniation, lipomas, Dercum's disease, lipodystrophy, buffalo hump lipodystrophy, dorsocervical fat, visceral adiposity, breast enlargement, hyperadiposity, diffused body fat around trunk and arms, and fat deposits associated with cellulite. In preferred embodiments, the above method does not include performing surgery on said subject.

**[0062]** In one aspect, the present invention relates to methods for reducing the appearance of a skin condition in a skin region of a subject. Such methods comprise the step of: administering locally to said skin region a composition comprising: (i) a skin-tightening effective amount of one or more pharmacologically active detergents, or bile acid(s) and/or bile salt(s), or deoxycholic acid or a salt thereof, or sodium deoxycholate, (ii) a pharmaceutical, veterinary, or cosmetic excipient, and (iii) optionally a lipid. In some embodiments, the administering step involves delivering the compositions herein via a subcutaneous or transdermal injection. In some embodiments, the skin condition being treated or ameliorated is selected from the group consisting of: loose skin, skin aging, irregularities of the skin, and wrinkles. In some embodiments, the region of skin



being treated is under eye, under chin, under arm, buttock, cheek, brow, calf, back, thigh, ankle, or stomach.

**[0063]** In some embodiments, the compositions used for reducing the appearance of a skin condition in a skin region are formulated into a skin tightening solution. Such skin tightening solution can further comprise a second therapeutic agent selected from the group consisting of: anti-microbial agents, vasoconstrictors, anti-thrombotic agents, anti-coagulation agents, suds-depressants, anti-inflammatory agents, analgesics, dispersion agents, anti-dispersion agents, penetration enhancers, steroids, tranquilizers, muscle relaxants, and anti-diarrhea agents.

**[0064]** In preferred embodiments, the detergent comprises a bile acid selected from the group consisting of deoxycholic acid, cholic acid, chenodeoxycholic acid, 7- $\alpha$ -dehydroxylate chenodeoxycholic acid, lithocholic acid, ursodeoxycholic acid, dihydroxytaurine acid, trihydroxytaurine acid, and glycine conjugates of any of the above. In some embodiments, the detergent comprises a bile salt that includes a cation selected from the group consisting of sodium ( $\text{Na}^+$ ), potassium ( $\text{K}^+$ ), lithium ( $\text{Li}^+$ ), magnesium ( $\text{Mg}^{2+}$ ), calcium ( $\text{Ca}^{2+}$ ), barium ( $\text{Ba}^{2+}$ ), strontium ( $\text{Sr}^{2+}$ ), ammonium ( $\text{NH}_4^+$ ), primary, secondary and tertiary alkyl and aryl ammonium salts ( $\text{NR}^{10}\text{H}_3^+$ ,  $\text{N}(\text{R}^{10})_2\text{H}_2^+$  and  $\text{N}(\text{R}^{10})_3\text{H}^+$ , respectively, wherein  $\text{R}^{10}$  is independently alkyl or aryl). In some embodiments, the detergent comprises a bile salt with a cation that is an alkali metal or an alkaline earth metal. Preferably, the alkali metal is sodium ( $\text{Na}^+$ ), potassium ( $\text{K}^+$ ), or lithium ( $\text{Li}^+$ ) and the alkaline earth metal is magnesium ( $\text{Mg}^{2+}$ ), calcium ( $\text{Ca}^{2+}$ ), barium ( $\text{Ba}^{2+}$ ), or strontium ( $\text{Sr}^{2+}$ ). More preferably, the bile salt is sodium deoxycholate.

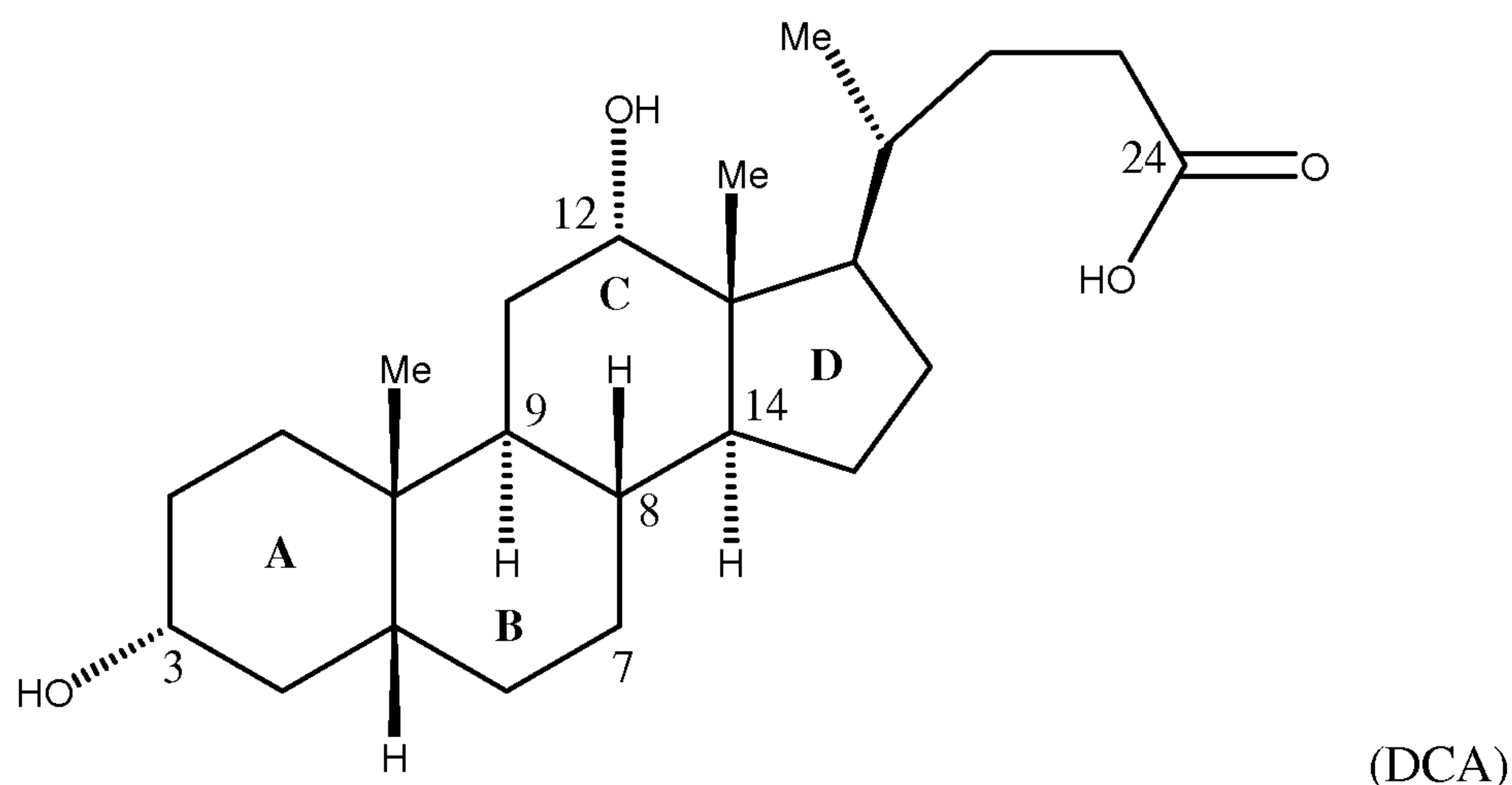
**[0065]** Sodium deoxycholate is a naturally produced bile salt that solubilizes dietary lipids in the digestive tract. It is produced *in vivo* via a complex biosynthetic route utilizing cholesterol as the starting material and involving both human and bacterial enzymes. The primary function of deoxycholate is to assist in the digestive process by solubilizing dietary lipids to facilitate absorption. In the body, deoxycholate biosynthesis begins with the enzymatic oxidation, isomerization, and reduction of cholesterol in the liver to form cholic acid, a bile acid structurally similar to its cholesterol parent (Stryer L, *Chapter 27: Biosynthesis of Membrane Lipids and Steroids*, in *Biochemistry*, 1995, W. H. Freeman and Company: New York. p. 691-707). In the liver, cholic acid is then chemically linked to one of two amino acids (taurine or glycine) to form the 'conjugated'

cholic acids (*i.e.*, L-glycocholate and taurocholate). These conjugated cholic acids are then stored in the gall bladder until food consumption. After food consumption, bile solution is released from the gall bladder into the intestine, where the conjugated cholic acid molecules are subject to two additional chemical modifications mediated by enzymes produced by intestinal microflora (Ridlon J.M., Kang D.J. and Hylemon P.B., *Bile salt biotransformations by human intestinal bacteria*, J. Lipid Res., 2006, **47**(2): p. 241-59). First, conjugated cholic acid is dehydroxylated to form conjugated deoxycholate. Conjugated deoxycholate is then deconjugated to form free deoxycholate, which participates, along with the other bile acids, in the solubilization of dietary lipids. Because deoxycholate is downstream from cholic acid synthesis, cholic acid may be an impurity present in natural sources of deoxycholate.

**[0066]** Deoxycholate is soluble to 333 mg/mL in water, sparingly soluble in alcohol, and is even less soluble in acetone and glacial acetic acid. Reversible formation of micelles may occur with sodium deoxycholate concentrations above the critical micelle concentrations of approximately 2.4 mg/mL and neutral pH (Matsuoka K, M.Y., *Micelle formation of sodium deoxycholate and sodium ursodeoxycholate (part 1)*, Biochim. Biophys. Acta., 2002, **1580**(2-3): p. 189-99). At concentrations above the critical micelle concentration of 2.4 mg/mL, deoxycholate will form micelles and has the ability to solubilize cells, lipids, and proteins. At lower concentrations such as 0.4 mg/mL (comparable to the fasting state), deoxycholate is 98% bound to albumin (Roda A. et al., *Quantitative aspects of the interaction of bile acids with human serum albumin*, J. Lipid Res., 1982, **23**(3): p. 490-5) in the presence of 26 mg/mL of albumin (which is close to the serum physiological concentration of 35-50 mg/mL).



[0067] The preferred embodiments are directed to deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof and the related compositions and methods, wherein deoxycholic acid (DCA) is:



wherein said compound is not isolated from a mammalian or microbial organism naturally producing DCA.

[0068] Other preferred embodiments also are directed to stereoisomers of DCA and pharmaceutically acceptable salts thereof and to intermediates in the synthesis of the DCA and its stereoisomers and salts and the related compositions and methods.

[0069] Methods for complete chemical synthesis of bile acid pharmaceutical compositions, and useful intermediates, are now provided.

[0070] These following descriptions and examples provide an alternative to the extraction of DCA from mammalian or microbial organisms that naturally produce this compound. Synthetic routes 1-6 are contemplated for use in the present invention to synthesize deoxycholic acid (DCA). Synthetic route 1B and Examples 1-11 show the synthesis of DCA from hydrocortisone.

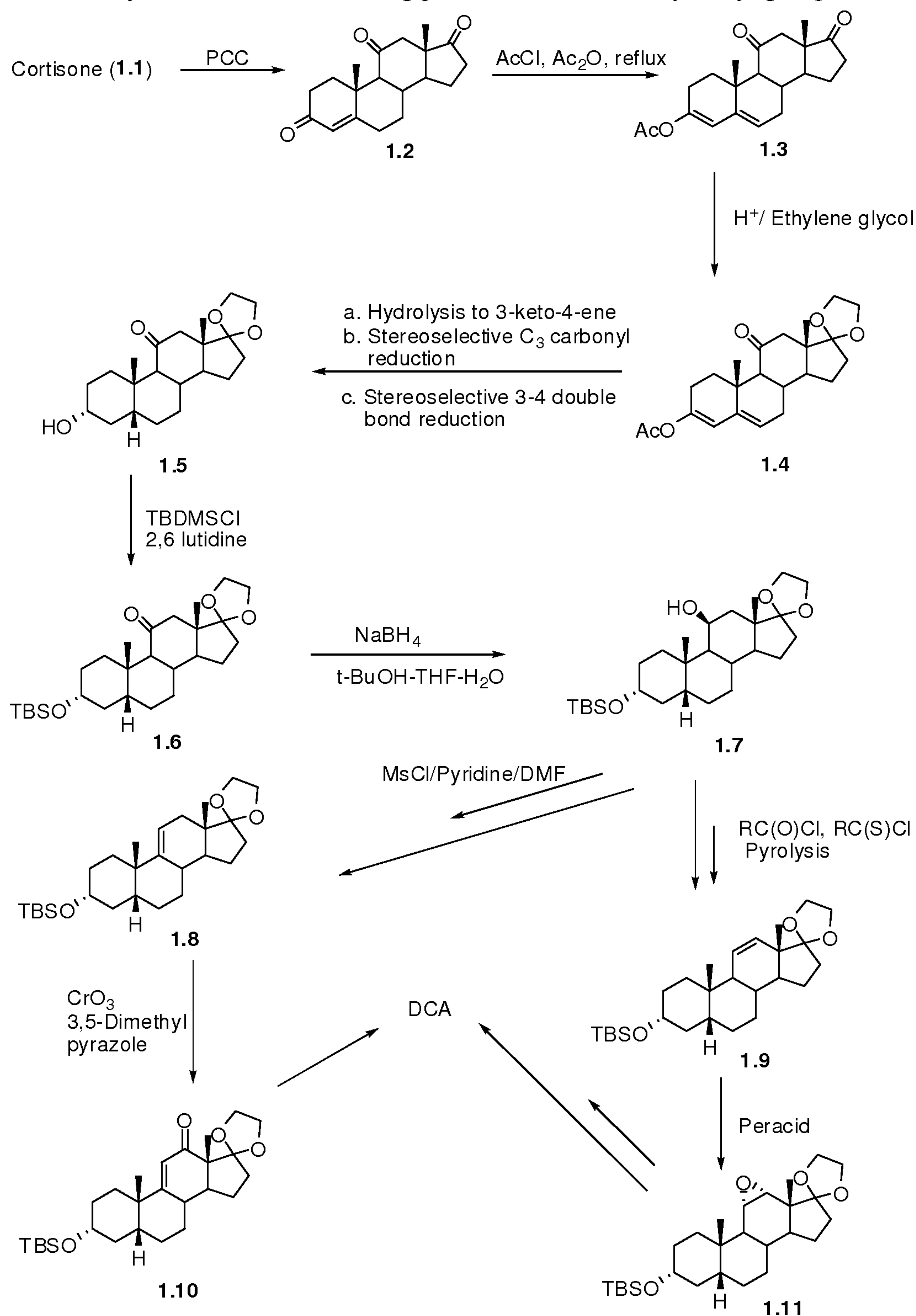
### 1. Synthetic route #1A from Adrenosterone, via 9(11)-ene or 11,12-ene

[0071] Cortisone (Compound 1.1) of Scheme 1A (below) is widely available as a fully synthetic material. It can be efficiently cleaved to form the C<sub>17</sub> ketone compound using pyridinium chlorochromate (PCC). This cleavage to adrenosterone (Compound 1.2) can also be achieved using HIO<sub>4</sub> or sodium bismuthate (NaBiO<sub>3</sub>). The reaction that converts Compound 1.2 to Compound 1.3 is a known chemical process. Conversion of Compound 1.3 into Compound 1.4 involves monoketalization. Subsequent steps are regeneration of

the 3-keto-4-ene, selective reduction of the 4,5-ene ( $\text{H}_2/\text{Pt}/\text{DMF}$ ) to yield the  $\text{C}_5$   $\beta$ -configuration and selective reduction of the  $\text{C}_3$  carbonyl group to the desired  $3\alpha$ -configuration to yield compound **1.5**. The addition of a protecting group in converting Compound **1.5** into Compound **1.6** and subsequent reduction of the product yields the  $\text{C}_{11}$   $\beta$ -ol (axial configuration), i.e., compound **1.7**, which is suitable for regioselective dehydration to the key 9(11)-ene (i.e., conversion of Compound **1.7** into Compound **1.8**).

[0072] The synthetic scheme bifurcates here, in that Compound **1.7** can be used as the starting material for conversion to either Compound **1.8** or Compound **1.9**. The elimination reaction used to convert Compound **1.7** into Compound **1.8** is regioselective because of the trans diaxial relationship between the  $\text{C}_{11}$  hydroxyl group and  $\text{C}_9$  hydrogen atom. The alternative mode of elimination to yield the isomeric  $\text{C}_{11}$ - $\text{C}_{12}$  olefin of compound **1.9** is likewise regioselective involving *cis*-thermal elimination (i.e., conversion of Compound **1.7** to Compound **1.9**).



Scheme 1A. Synthesis of the two C-ring precursors of the C<sub>12</sub> hydroxyl group of DCA

[0073] Allylic oxidation of Compound 1.8 (via treatment with CrO<sub>3</sub> and 3,5 dimethyl pyrazole) yields the enone-containing Compound 1.10. Peracid oxidation of Compound 1.9 proceeds stereoselectively from the alpha-face of the steroid to yield the C<sub>11-12</sub> epoxide Compound 1.11 (see Scheme 1A above). These chemical transformations yield

the two key precursors of the C<sub>12</sub> hydroxyl group functionality, namely, Compound **1.10** and Compound **2.1** (Schemes 1A and 2).

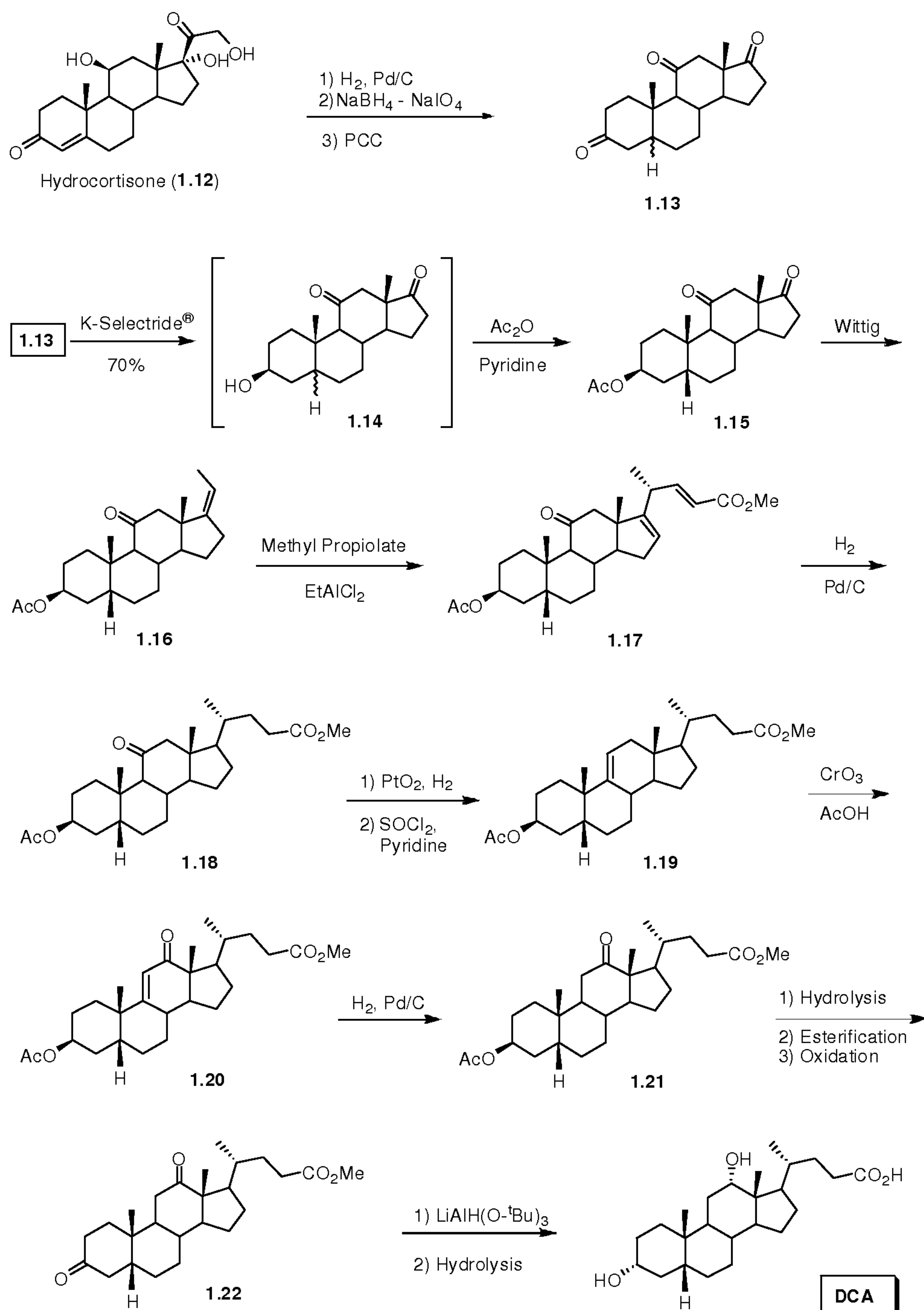
[0074] One of the skill in the art will appreciate that the above cortisone route can be modified to begin instead with hydrocortisone, which has the same carbon skeleton and the same relative placement of oxygen atoms, with hydrocortisone differing from cortisone only in the oxidation state of the C-11 oxygen bearing carbon atom.

Hydrocortisone is commercially available and various synthesis of this compound are known (Szczepara et. al. Nature Biotechnology Vol. 21, Feb. 2003, 143-149) including a total chemical synthesis (Woodward R. B. et. al. 1952, J. Am. Chem. Soc. 74: 4223).

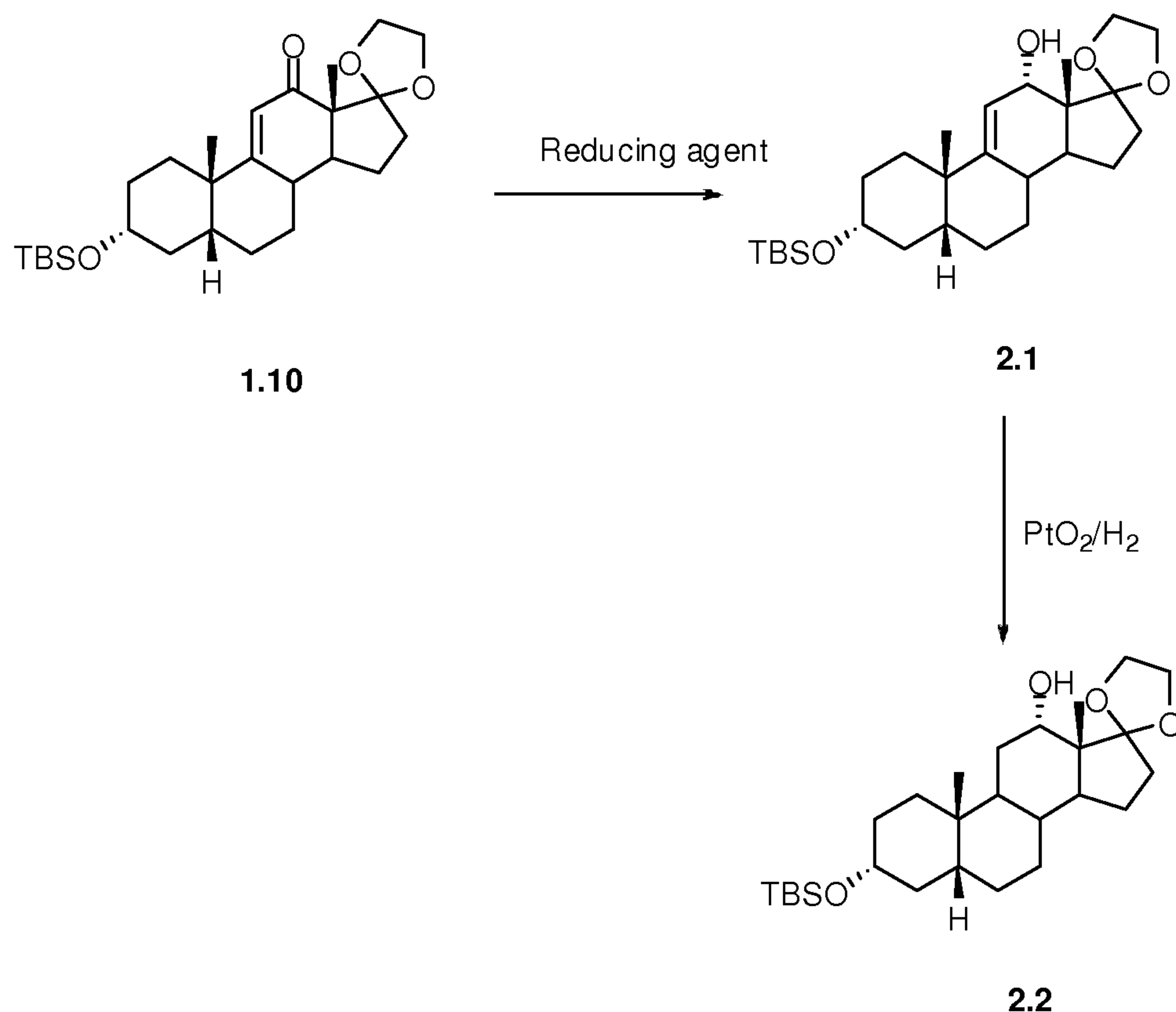
Ketone **1.13** is synthesized starting from hydrocortisone **1.12** (Scheme 1B) via hydrogenolysis of the  $\alpha,\beta$ -unsaturated double bond, followed by global ketone reduction using sodium borohydride to allow for 1,2-diol cleavage using NaIO<sub>4</sub>, thus forming the C<sub>17</sub> ketone on the D-ring of the steroidal ring system. Subsequent oxidation with pyridinium chlorochromate (PCC) yields **1.13**. Treatment of **1.13** with K-selectride<sup>®</sup> followed by acetylation with acetic anhydride/pyridine gives protected alcohol **1.15**. Subsequent olefination of **1.15** with a Wittig reagent provides alkene **1.16** that is then treated with methyl propiolate and ethyl aluminum dichloride to form diene **1.17**. Following hydrogenation of both double bonds, ketone **1.18** is reduced and the resulting alcohol intermediate is eliminated upon treatment with SOCl<sub>2</sub> in pyridine to give alkene **1.19**. Allylic oxidation of alkene **1.19** with CrO<sub>3</sub> and reduction of the double bond under hydrogenation conditions gives ketone **1.21**. Removal of the acetate protecting group and oxidation of the resulting alcohol gives diketone **1.22**. Reduction of **1.22** with LiAlH(O-<sup>t</sup>Bu)<sub>3</sub> and hydrolysis of the methyl ester yields DCA.



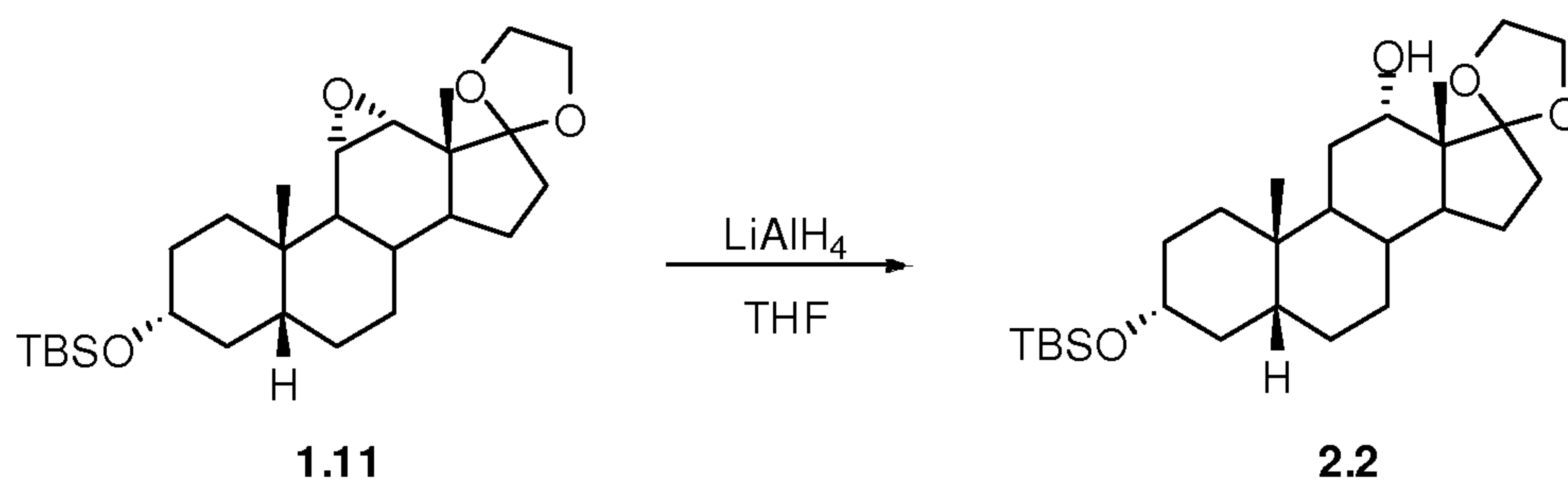
## Scheme 1B. Synthesis of DCA from Hydrocortisone



[0075] Further transformations of Compounds **1.10** and **2.1** are shown in Scheme 2. First, Compound **1.10** is modified to contain a properly functionalized C ring system identical to that of DCA (Scheme 2). Stereoselective reduction of the  $\text{C}_{12}$  carbonyl group yields Compound **2.1** and catalytic hydrogenation of the 9(11) double bond present in Compound **2.1** yields Compound **2.2**.

Scheme 2. Introduction of C<sub>12</sub> α-hydroxyl group using the allylic oxidation route

[0076] Scheme 3 presents the transformation of epoxide-containing Compound **1.11** to the analogous C<sub>12</sub> α-hydroxy steroid Compound **2.2** of Scheme 2.

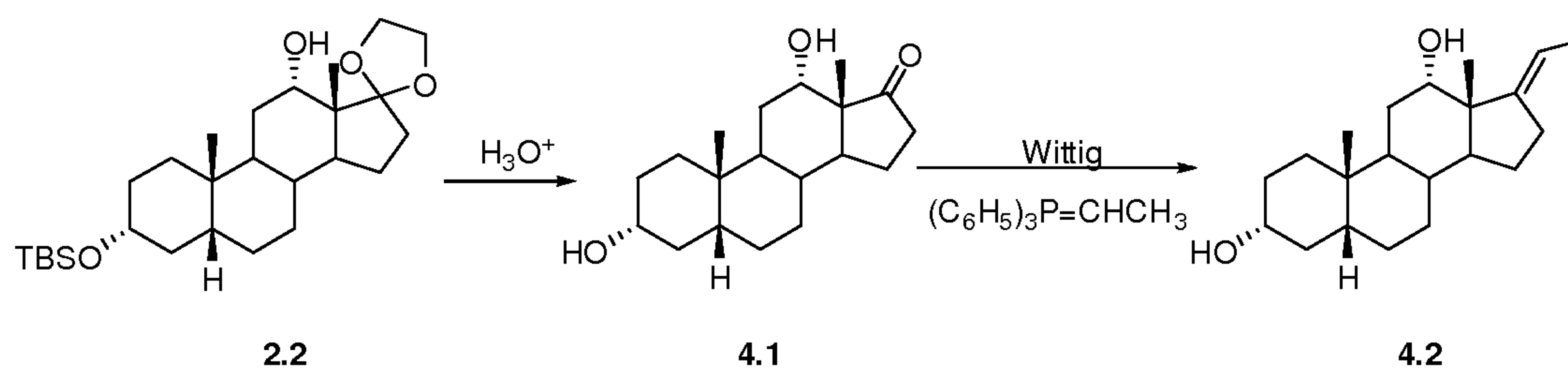
Scheme 3. Stereoselective Reduction of C<sub>11</sub>-C<sub>12</sub> epoxide .

[0077] As mentioned above in both of these routes common intermediate compound **2.2** is formed.

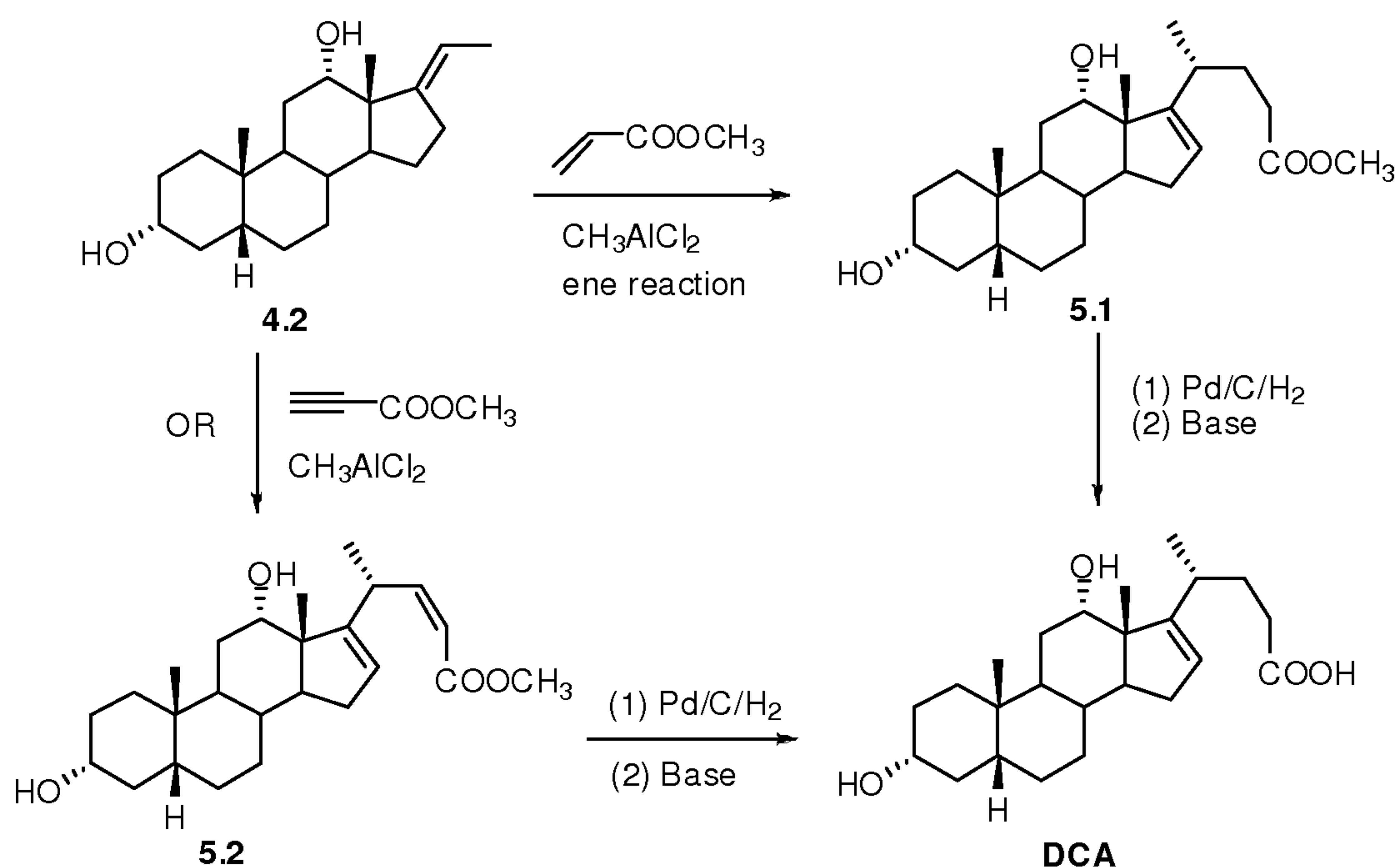
[0078] The next step in the synthesis of DCA is the modification of the D-ring present in Compound **2.2** such that it contains the carboxylic side chain substituted D ring of DCA (Scheme 4 and Scheme 5).



Scheme 4. Deprotection and Wittig Reaction



**[0079]** First the C<sub>17</sub> ketal and the C<sub>3</sub> silyl ether groups of Compound **2.2** are hydrolyzed. Then the Wittig reaction is performed to yield Compound **4.2**. Conversion of Compound **4.2** to Compound **5.1** is carried out via an *ene* reaction. Subsequent catalytic reduction of Compound **5.1** and hydrolysis of the ester yields DCA (Scheme 5).

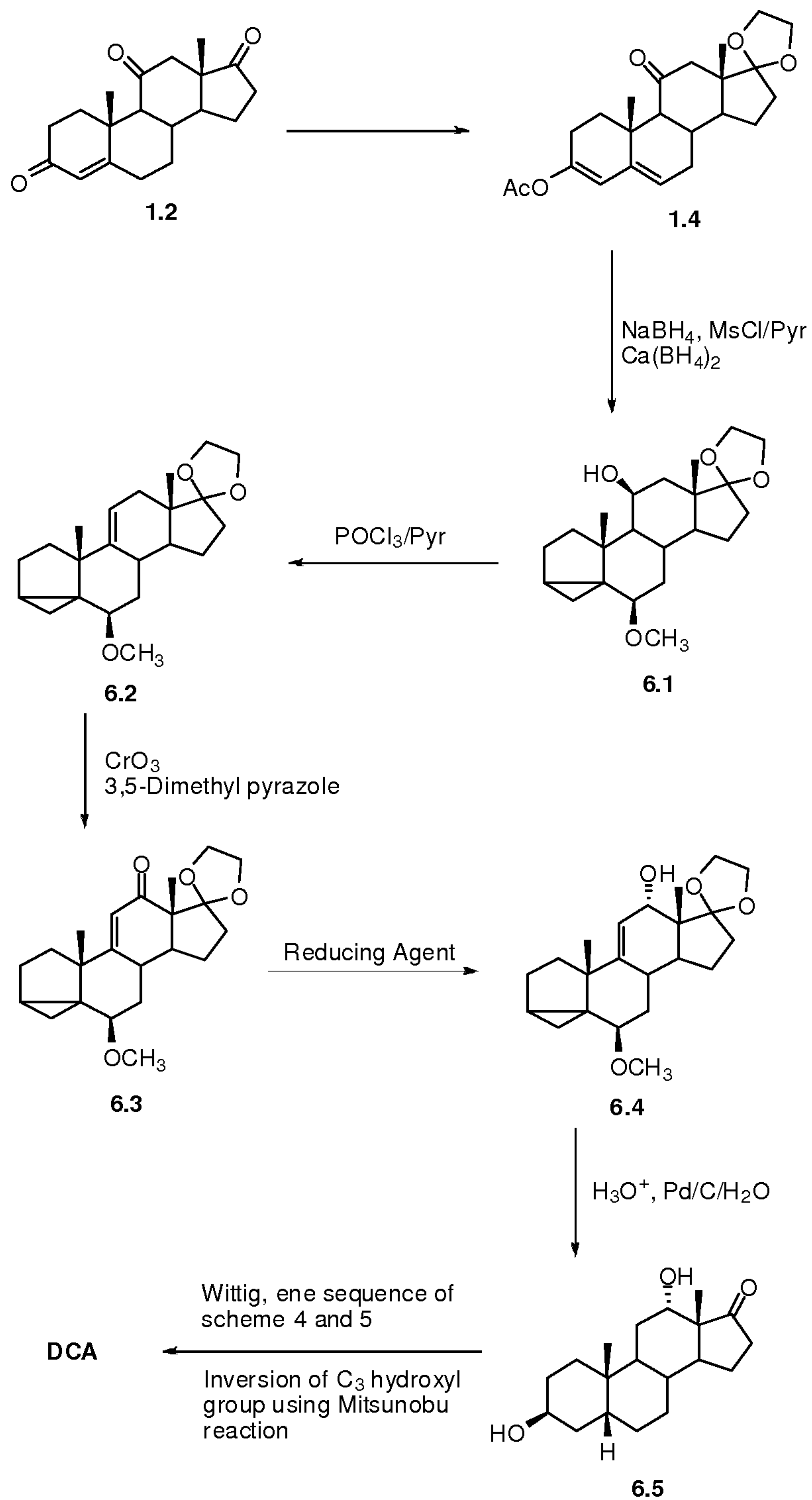
Scheme 5. *Ene* reaction and catalytic reduction for installation of DCA side chain

## 2. Synthetic route #2 from Cortisone via Adrenosterone (the i-Steroid, 3,5-cyclosterol route)

**[0080]** Selective ketalization of adrenosterone (Compound **1.2**, Scheme 6) at C<sub>17</sub>, borohydride reduction, mesylation, and buffered methanolysis yields the i-steroid (3,5-cyclosterol) containing Compound **6.1**. Compound **6.1** undergoes 9(11)-*ene* formation (conversion of Compound **6.1** to Compound **6.2**, Scheme 6) and allylic oxidation (conversion of Compound **6.2** to Compound **6.3**, Scheme 6) followed by carbonyl group reduction to yield Compound **6.4**. Hydrolysis of the i-sterol and hydrogenation yields

Compound **6.5**, which can be converted to DCA by synthetic methods presented above in Synthetic Route #1.

Scheme 6. Protection of A-B- ring system by formation of the i-steroid (i.e., 3,5-cyclosterol)



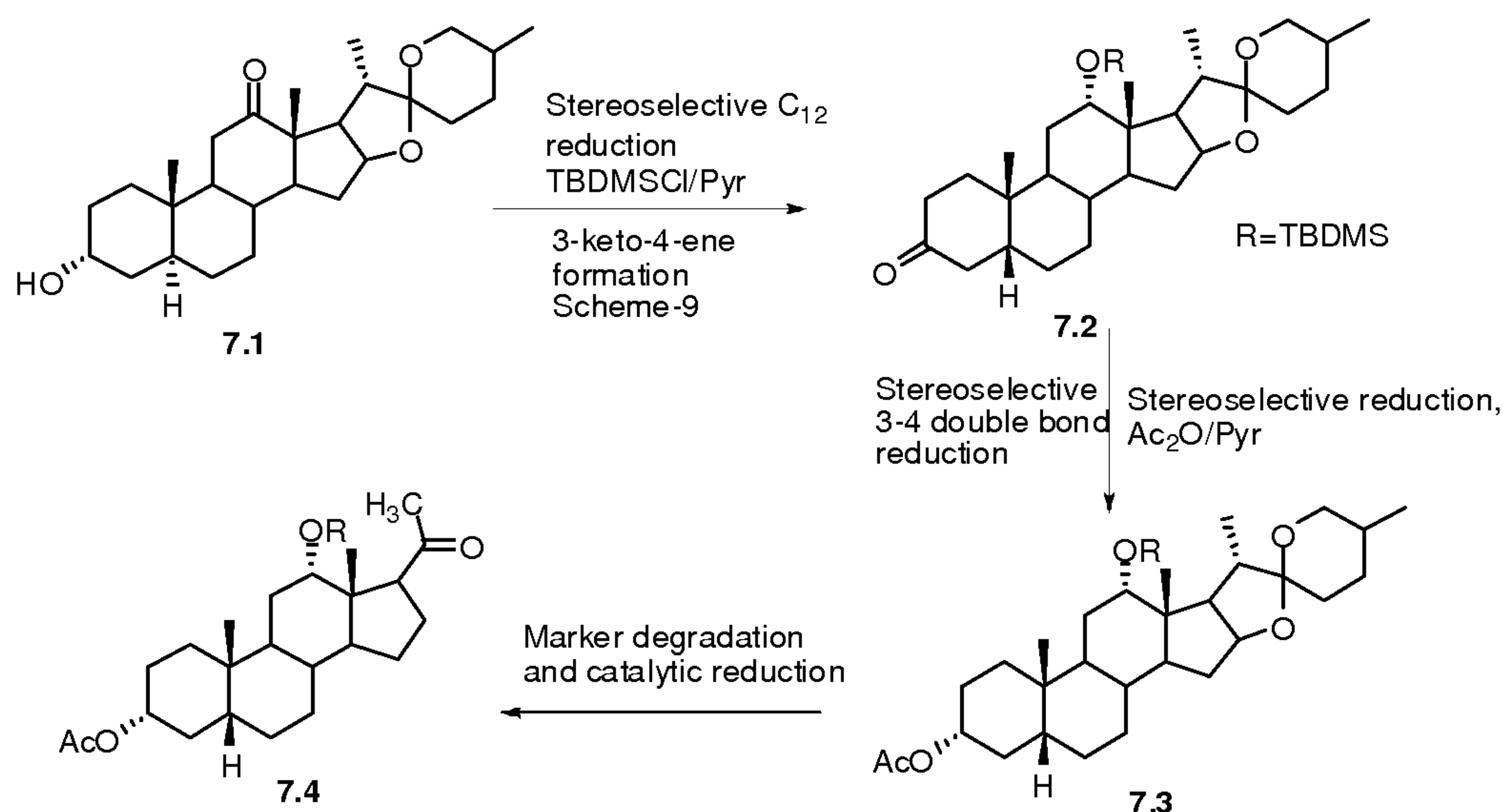


### 3. Synthetic route #3 from Hecogenin

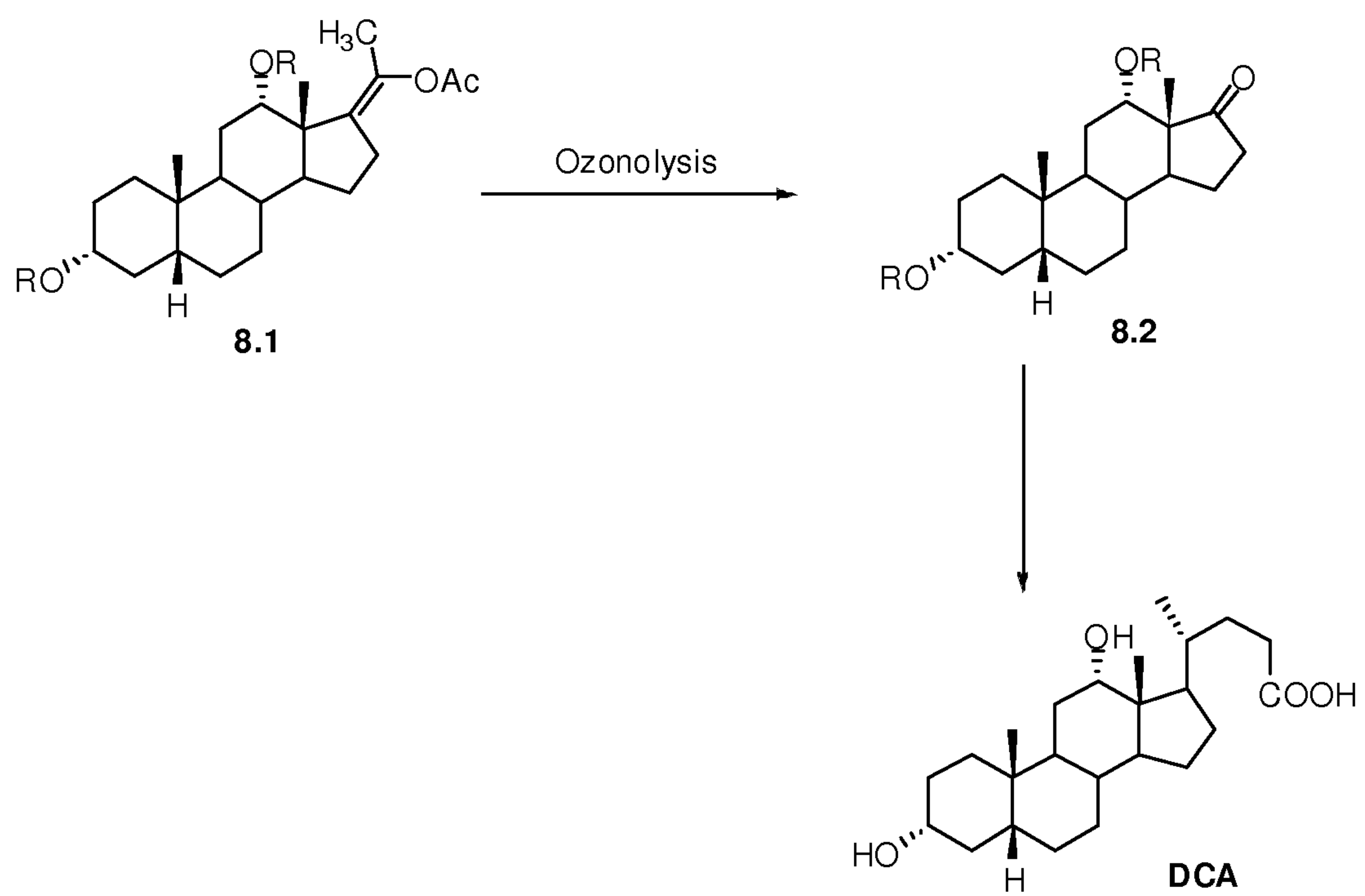
[0081] Hecogenin (Compound **7.1**, Scheme 7) is a plant sterol found abundantly in Mexican yams and other plants of the Agave species. The central advantage of hecogenin as a starting material for DCA synthesis is that it possesses a C<sub>12</sub> oxygen functionality as is present in DCA.

[0082] The first step in the synthetic route starting from hecogenin is the stereoselective reduction of the C<sub>12</sub> carbonyl group in hecogenin (Compound **7.1**) to the requisite C<sub>12</sub>- $\alpha$  configuration (conversion of Compound **7.1** to Compound **7.2**). Then the 3- $\beta$ -ol, 5 $\alpha$ -AB ring system is converted to the 3  $\alpha$ -ol, 5 $\beta$ -AB (Conversion of Compound **7.1**, to Compound **7.2**, to Compound **7.3**) ring system (Scheme 7). The well-known Marker degradation (Marker, R.E., Rohrmann, E., *Sterols. LXIX. Oxidation Products of Sarsasapogenin. Sarsasapogenoic Acid and Related Substances*, J. Am. Chem. Soc., 1939. 61(8): p. 2072 - 2077) follows the conversion of Compound **7.2** to Compound **7.3** to yield Compound **7.4**. Installation of the D-ring side chain (Scheme 8) into Compound **8.2** is achieved via methods shown in Schemes 4 and 5. The requisite C<sub>17</sub> ketone in Compound **8.2** is formed by ozonolysis of the enol acetate of Compound **8.1** (Scheme 8). DCA is then prepared from **8.2** in a similar manner as in Scheme 5 using the olefination and ene reaction sequence. Alternative routes starting from starting from hecogenin are shown in Schemes 9 and 10.

Scheme 7. C<sub>12</sub>-hydroxyl group introduction, AB ring modification and side chain cleavage

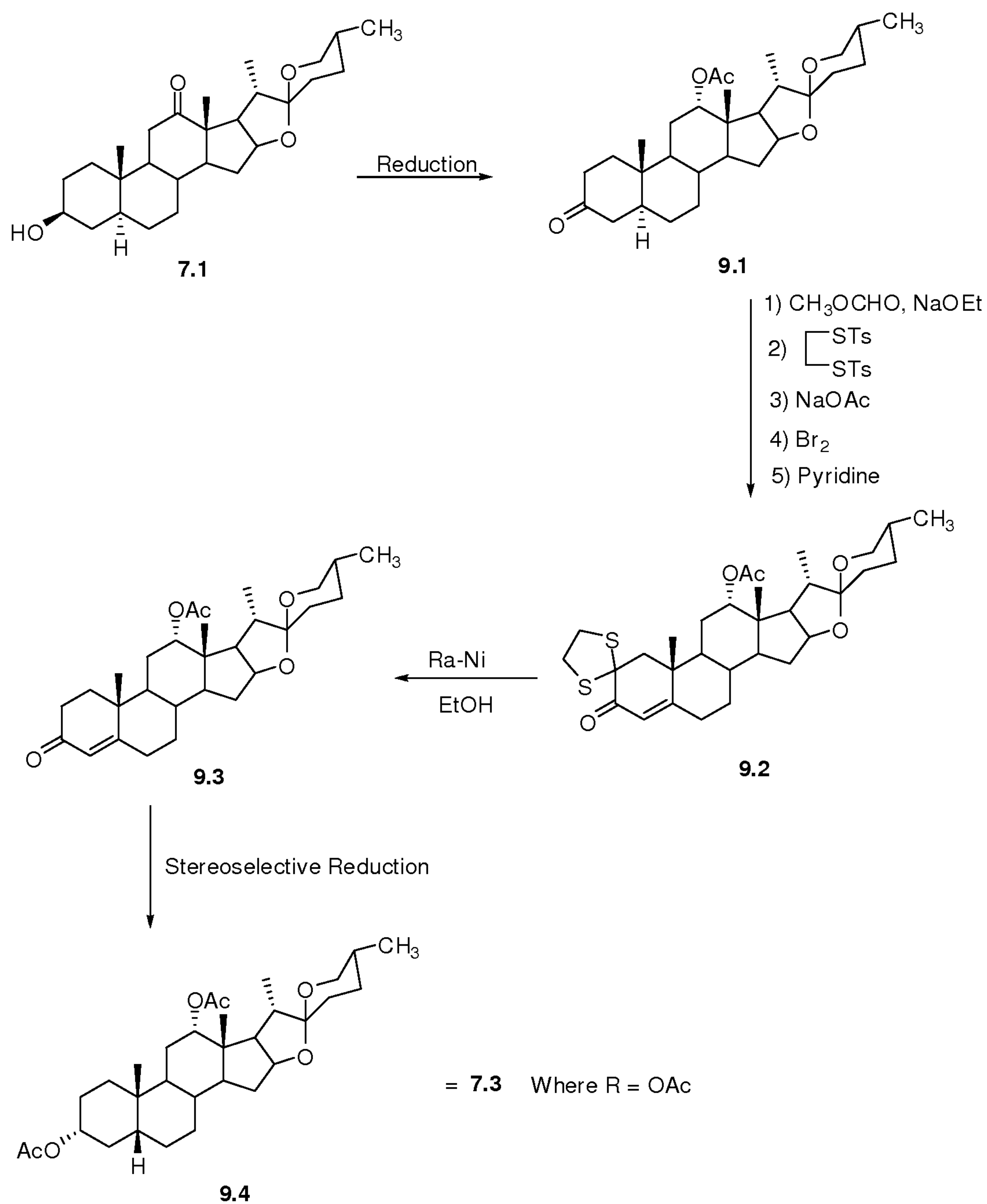


Scheme 8. Side-chain introduction

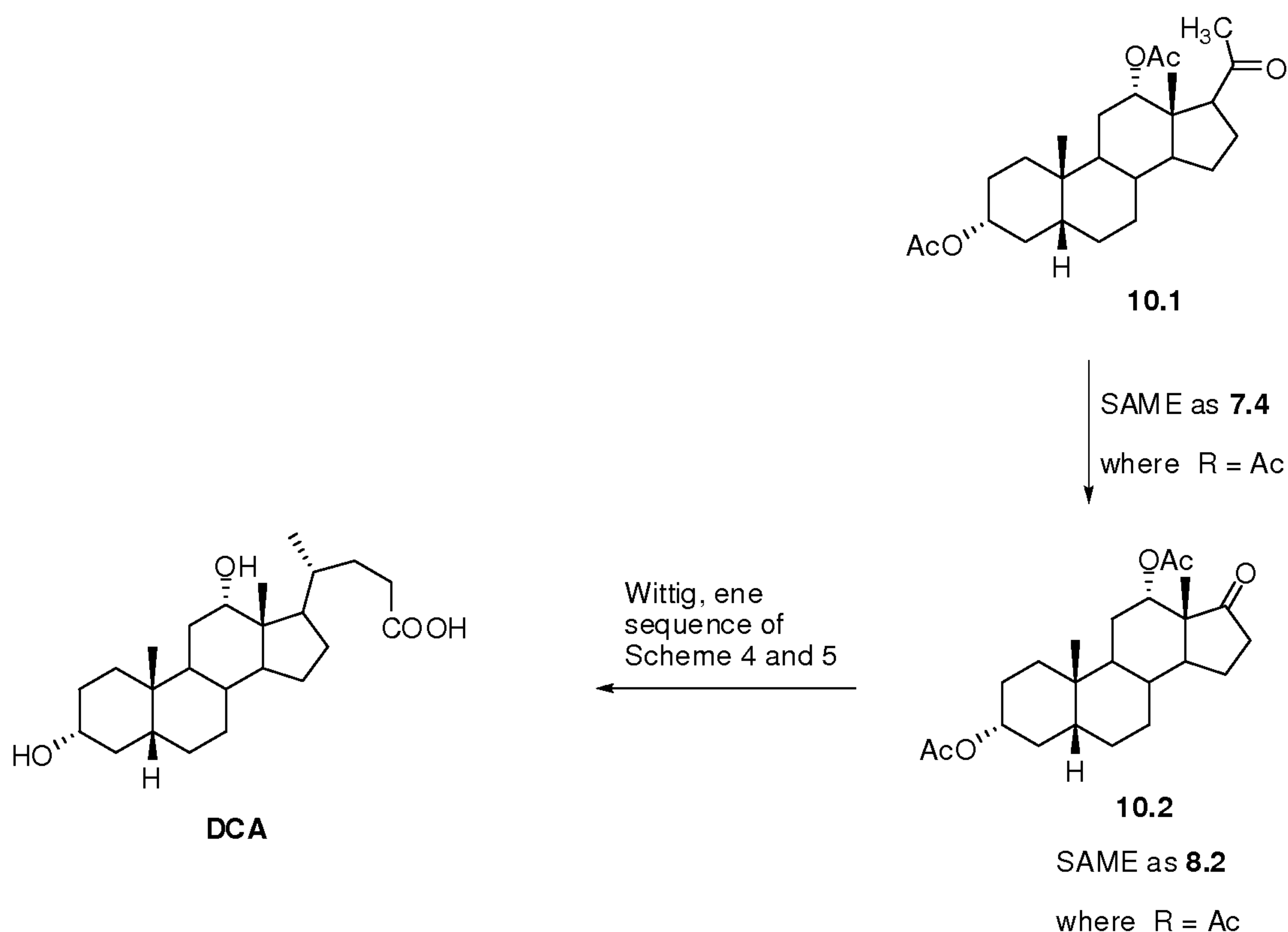




Scheme 9. Dithioethane route to 3-keto-4-ene



Scheme 10A. Formation DCA via the 3-keto-4-ene



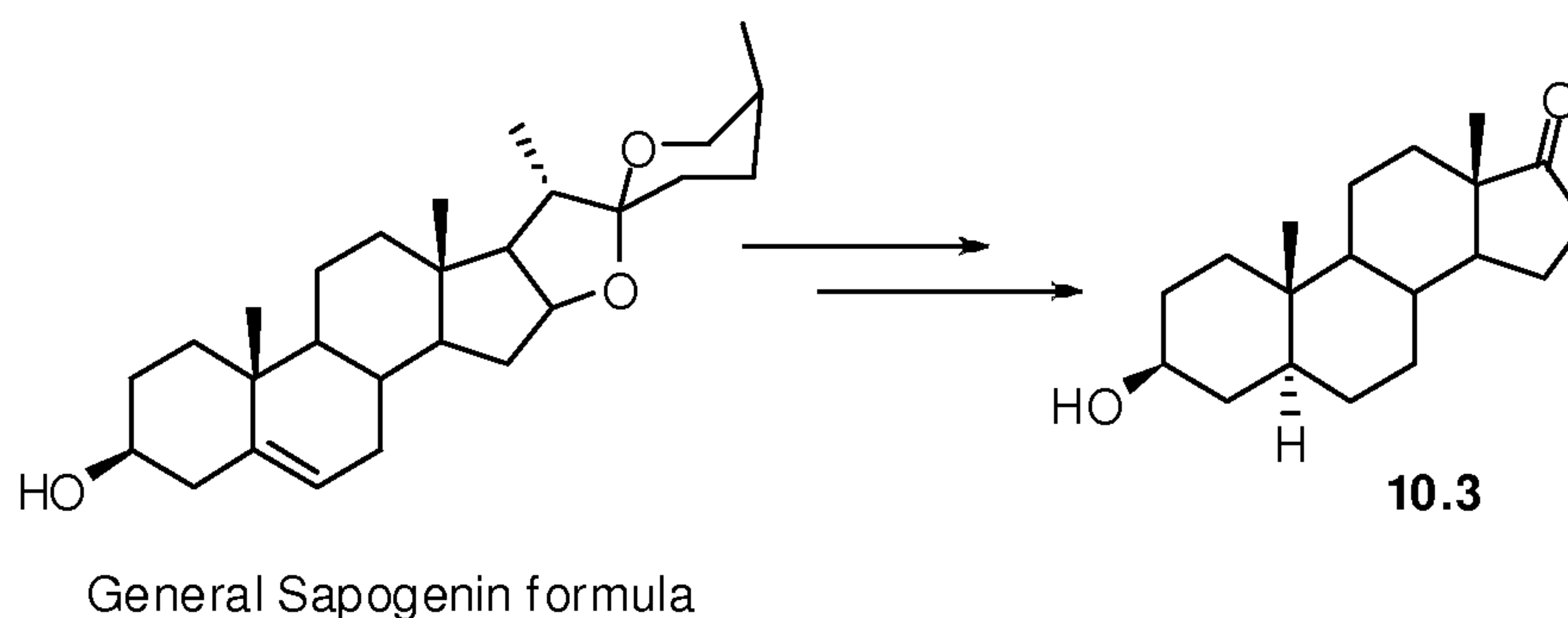
#### 4. Synthetic route #4 from Sapogenins

**[0083]** Sapogenins are derived from the hydrolysis of the saccharides and disaccharides attached to the C<sub>3</sub> hydroxyl group of the saponins (i.e., steroid glycosides). These are widely occurring plant products. Saponin occurs in nature as a spiroketal structure as shown below. Also Compound **10.3** can be formed from tigogenin, diosgenin, chlorogenin, smilagenin and hecogenin (Compound **7.1**). DCA could be synthesized from each of these, namely, tigogenin, diosgenin, chlorogenin, smilagenin and hecogenin (Compound **7.1**). (Y. Mazur, N. Danieli and Franz Sondheimer J. Am. Chem. Soc.; 82, 5809, 1960).

**[0084]** Since we could synthesize DCA from hecogenin, we recognize that any of the above sapogenins could, likewise, serve as a starting material for DCA synthesis.



Scheme 10B. Synthetic route from saponin.

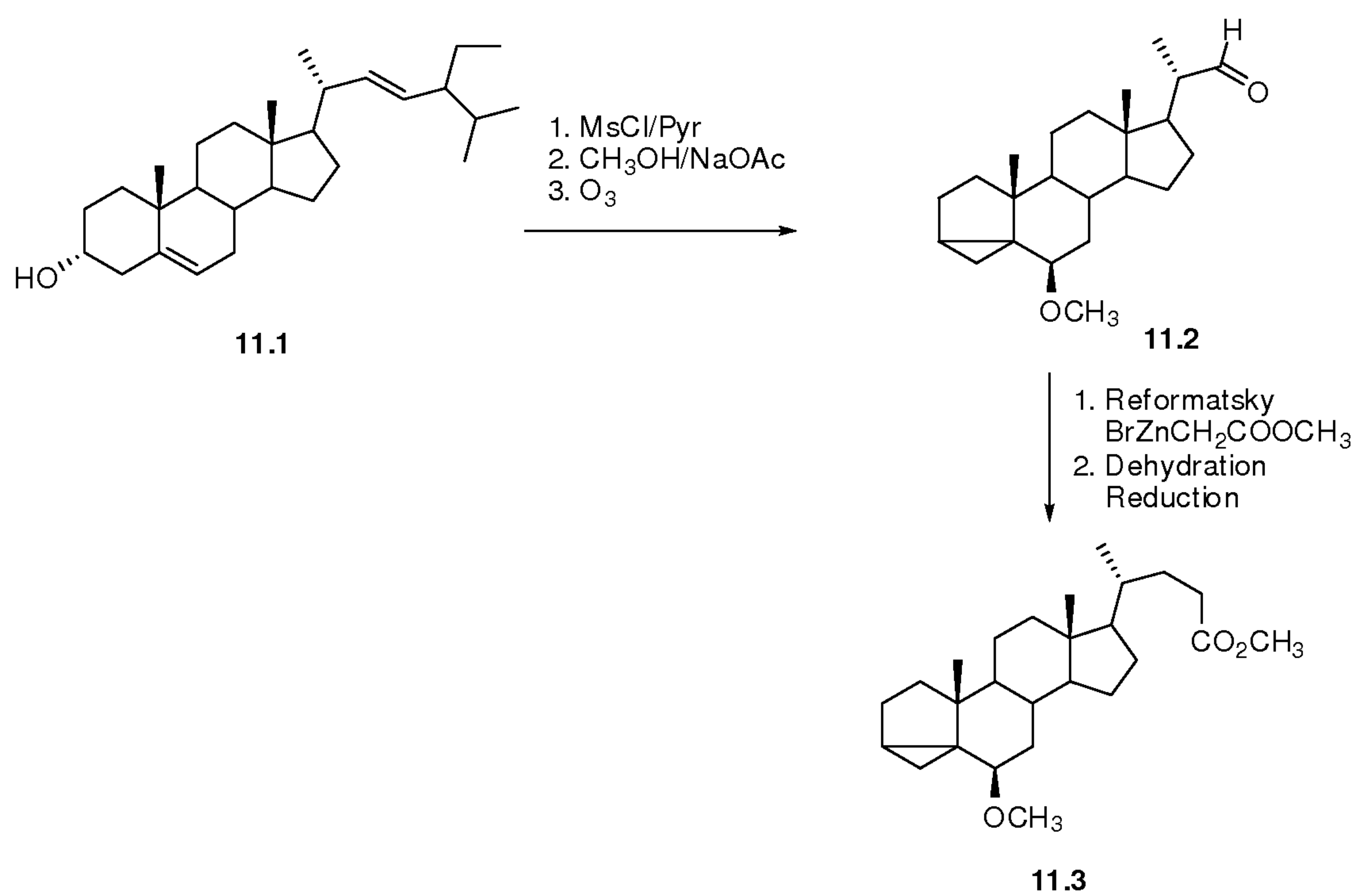


### 5. Synthetic route #5 from Stigmasterol

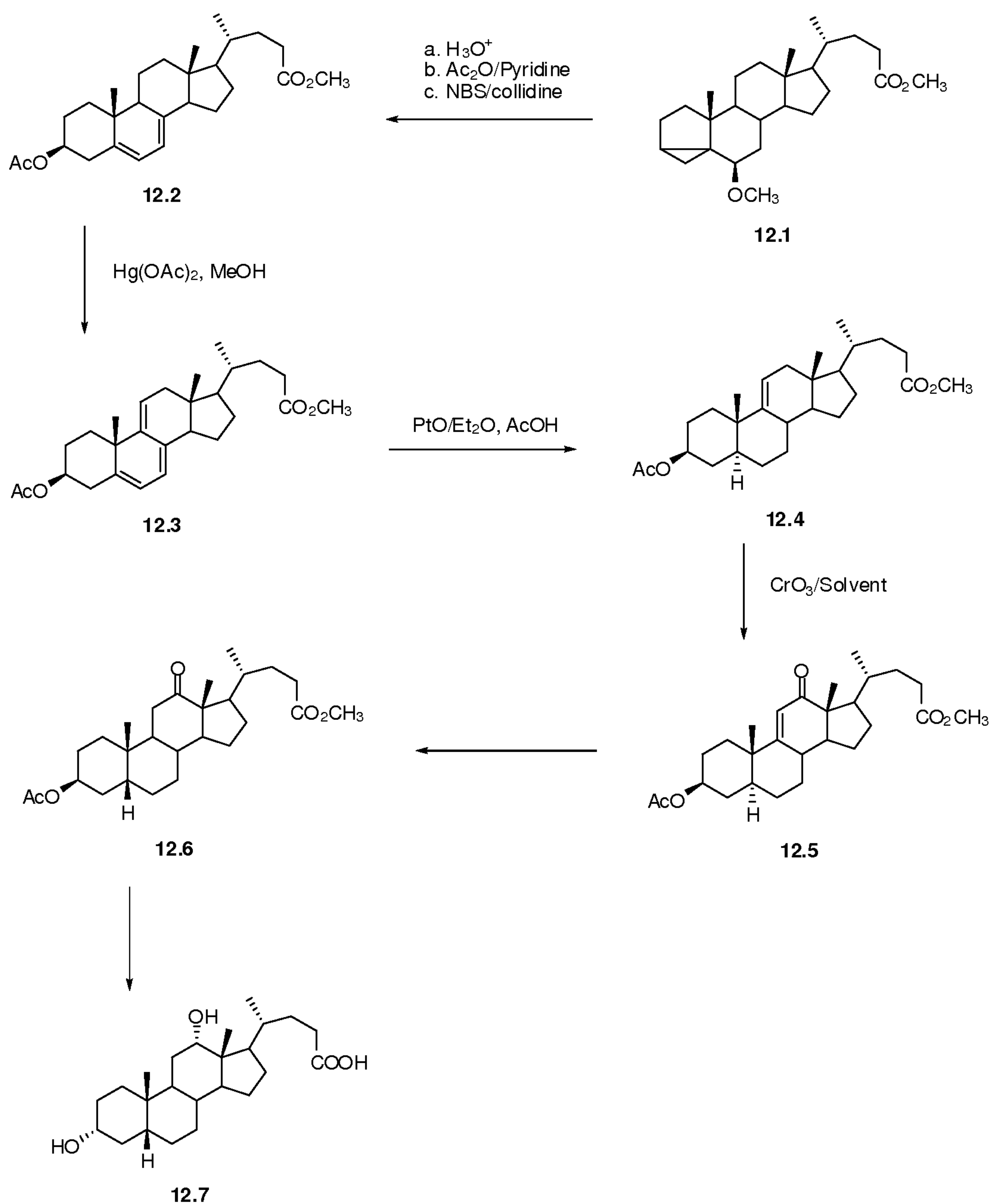
**[0085]** Stigmasterol (Compound **11.1**) is a widely available plant sterol. As a starting material for the DCA synthesis, it has an advantage in that it contains a functionalized AB ring system and a readily cleavable side chain moiety. It has the disadvantage in that it lacks the required functionality in the C ring essential for DCA synthesis.

**[0086]** In this synthetic route, the stigmasterol (Compound **11.1**) AB ring is protected by i-steroid formation followed by ozonolysis to yield a side-chain at C<sub>17</sub> installed and reduced to the C<sub>24</sub>-ol as a masked form of the carboxyl group (Scheme 11). The subsequent steps generate an allylic position at C<sub>12</sub> (Scheme 12). The B-ring diene formation and mercuric acetate oxidation are known processes and catalytic reduction of the B ring system yields an intermediate common with previous routes described above. However, in contrast to other routes, the side chain is already present. Allylic oxidation (conversion of Compound **12.4** to Compound **1.20**) and stereoselective reduction (conversion of Compound **1.20** to Compound **1.21**) followed by previously discussed steps, yields a product which is converted to DCA.

Scheme 11. Ozonolysis of i-steroid and side chain reduction



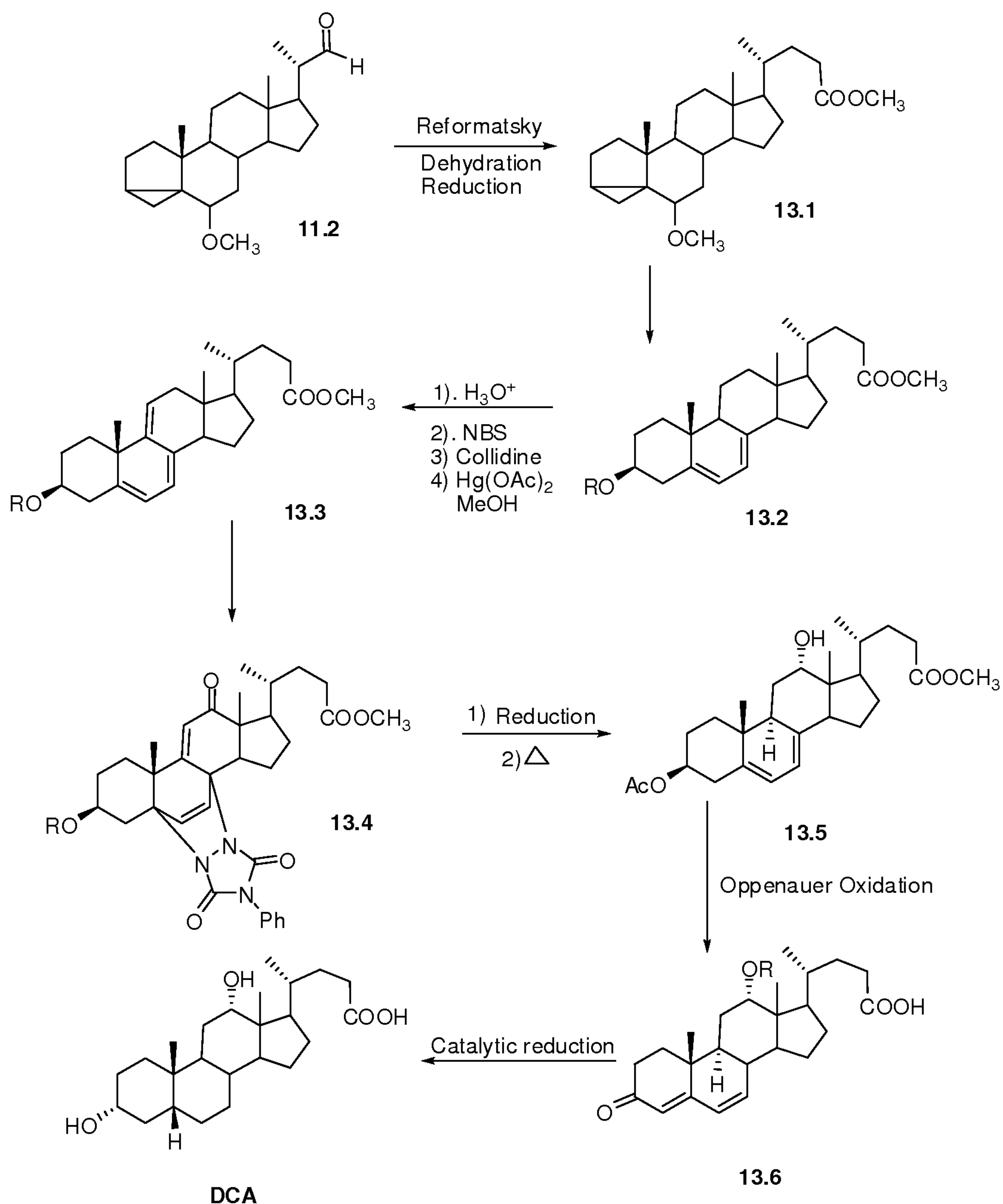
Scheme 12. Triene formation and allylic oxidation



**[0087]** A variation of the stigmasterol route uses the Diels-Alder protection of the B-ring diene. This is advantageous because it isolates the 9(11) double bond to prevent possible interference during the allylic oxidation steps (Scheme 13).



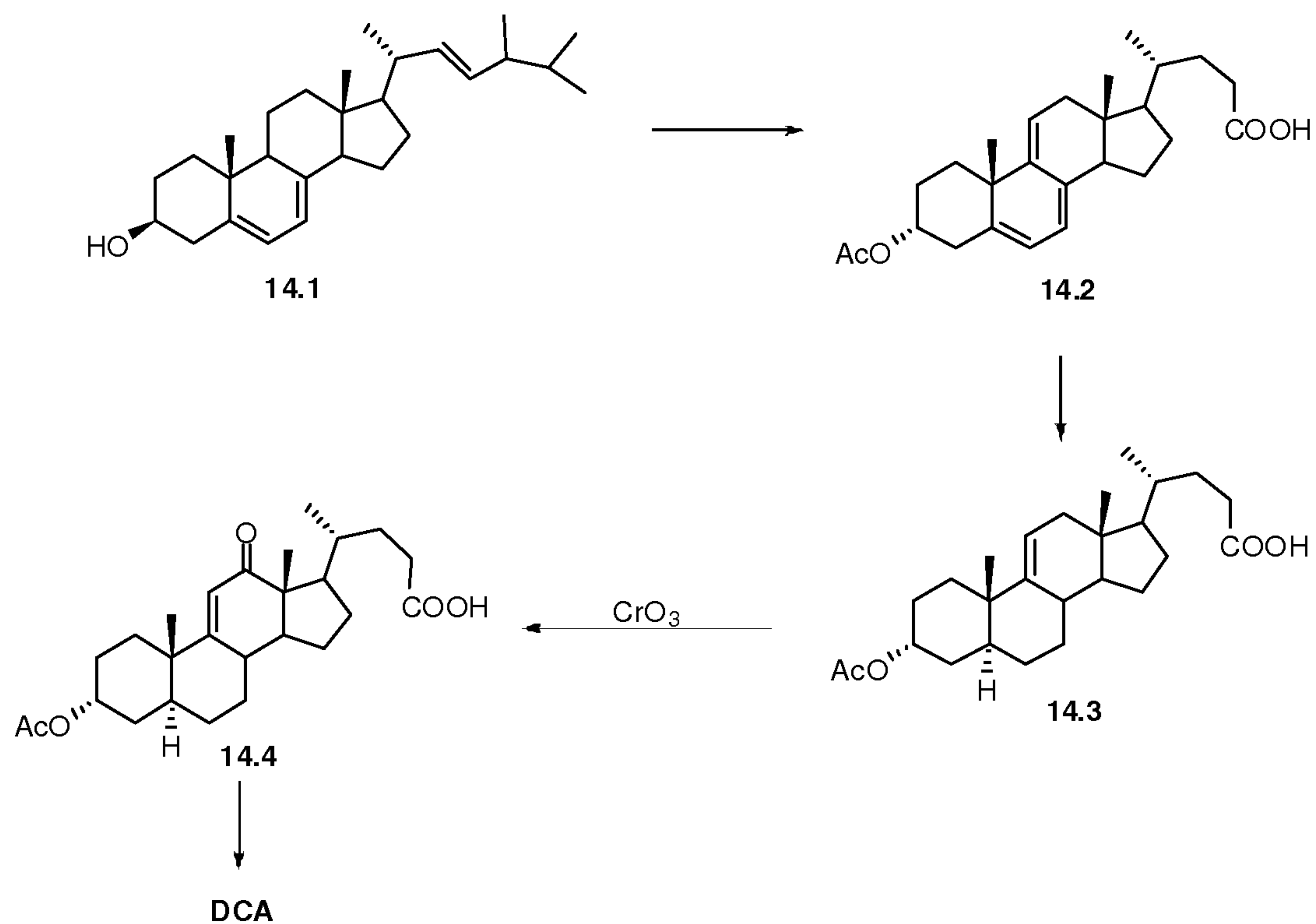
Scheme 13. Triene Formation and Diels-Alder Protection of the Ring B Diene.



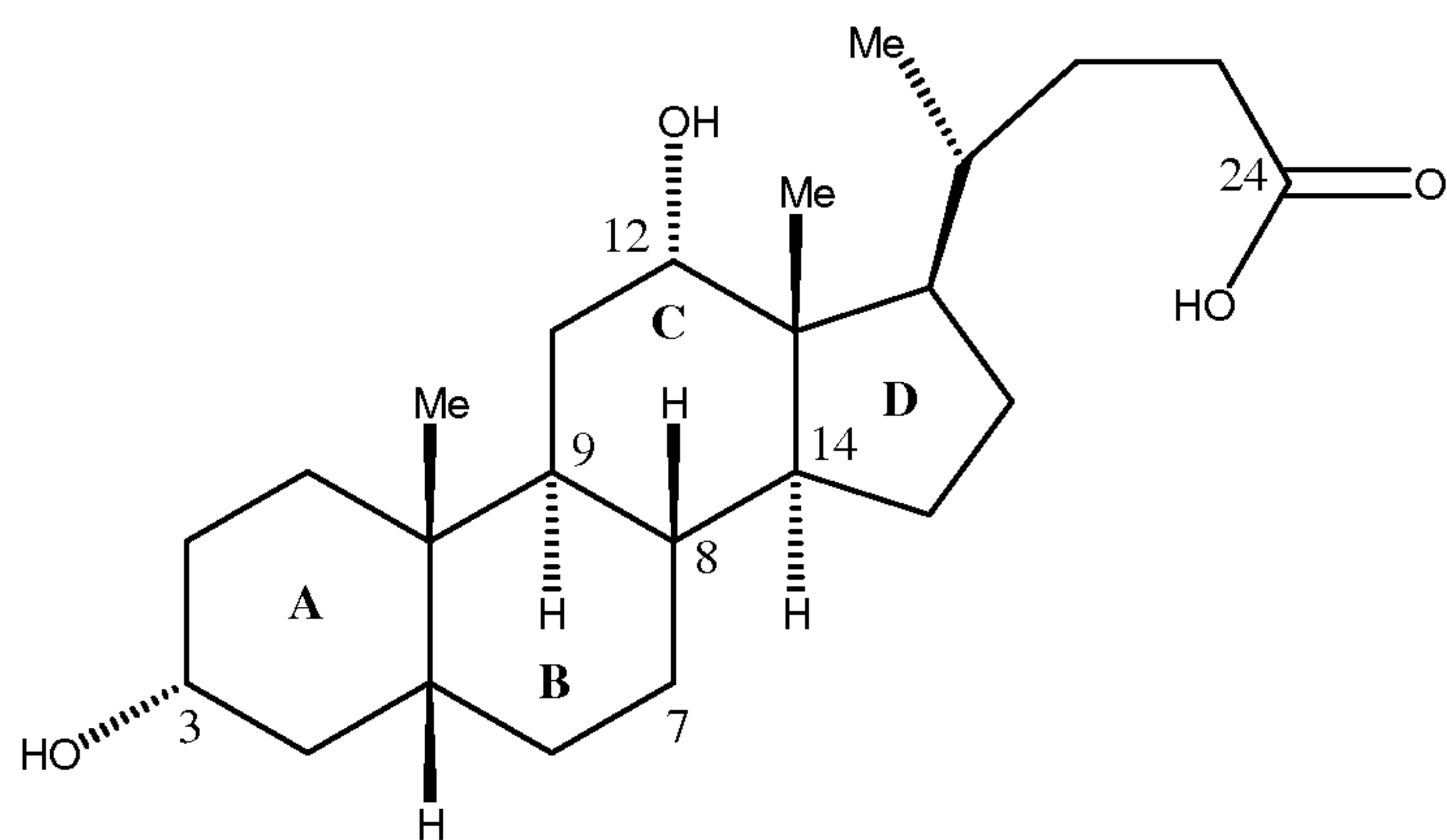
## 6. Synthetic route #6 from Ergosterol

[0088] Ergosterol (Compound 14.1) is a readily available starting material and can be used to prepare DCA by adaptation of the procedures set forth in this application. Allylic oxidation offers a facile route to  $\text{C}_{12}$  oxygen functionality (Scheme 14). This route has the advantage of starting with the ring B diene. It is convergent with the stigmasterol route.

Scheme 14. Triene formation from Ergosterol

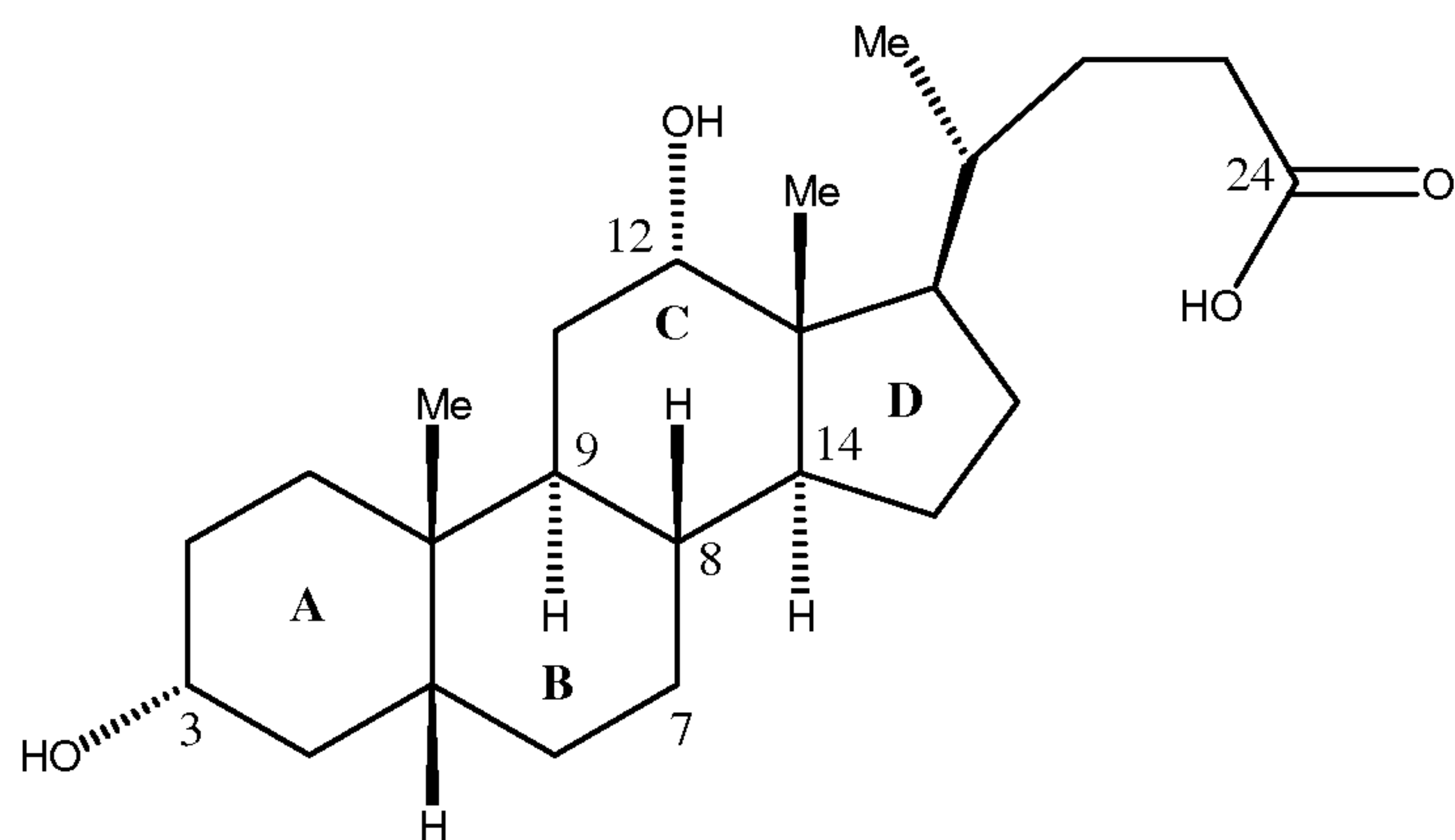


[0089] Another embodiment provides for a method for removal of fat deposits from selected locations in a mammal comprising administering to the mammal in need thereof a therapeutically effective amount of a compound that is DCA or pharmaceutically acceptable a salt thereof,



wherein said compound is not isolated from a mammalian or microbial organism naturally producing DCA.

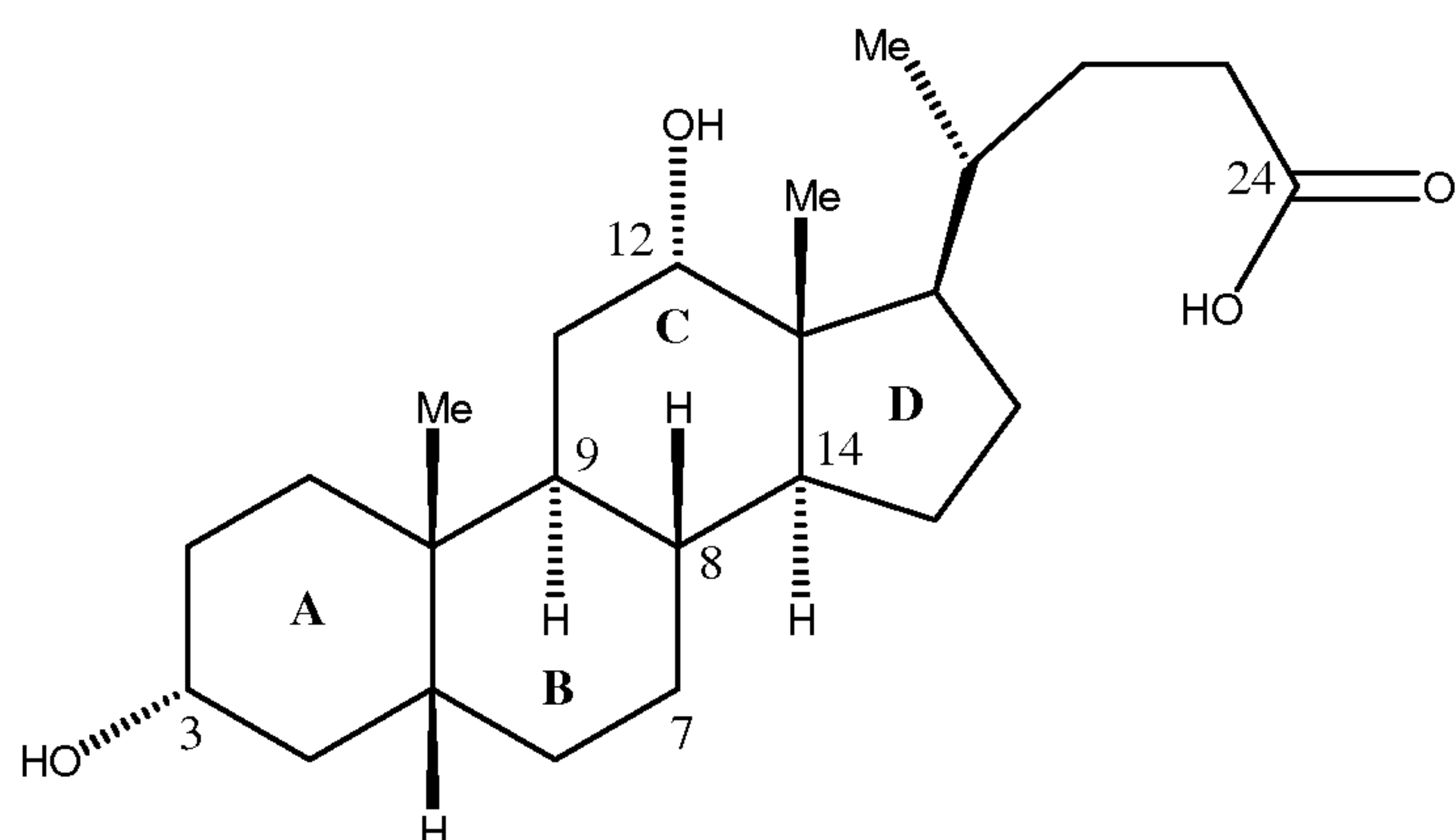
[0090] Another embodiment provides for a method of emulsifying fat in a mammal comprising administering to the mammal in need thereof a therapeutically effective amount of a compound that is DCA or pharmaceutically acceptable a salt thereof,



(DCA),

wherein said compound is not isolated from a mammalian or microbial organism naturally producing DCA.

[0091] Another embodiment provides for a method of solubilizing phosphatidylcholine comprising mixing phosphatidylcholine and effective amount of a compound that is DCA or pharmaceutically acceptable a salt thereof,



(DCA),

wherein said compound is not isolated from a mammalian or microbial organism naturally producing DCA.

[0092] In some embodiments, the DCA or pharmaceutically acceptable a salt thereof is as described herein.

[0093] Another aspect of the invention relates to mixing adipo-ablative bile acids, such as, deoxycholic acid (DCA) with agents that kill fat cells. In one aspect, this invention contemplates a means to enhance the aesthetic effects of deoxycholate injections by mixing into the deoxycholate injectate a molecule that causes fat to die by an orthogonal



mechanism. Examples of such candidate molecules include, but are not limited to, neuropeptide Y (NPY) antagonists and fat selective pro-apoptotic peptides. Since both fat cell killing and skin tightening may be required to mediate the desired effects, the effects of an agent with fat killing ability and potent skin tightening effects (such as deoxycholate) can be enhanced via the addition of a molecule with potent fat cell killing effects. Additionally, molecules that require access to the vasculature to kill (such as certain pro-apoptotic peptides that bind to proteins expressed on the luminal side of capillaries) can gain access to these proteins because deoxycholate may cause vascular leakage. Thus, such agents can be synergistic with deoxycholate potentially creating a more potent means to mediate body contouring in fewer therapeutic sessions.

**[0094]** Examples of NPY antagonists include, but are not limited to, NPY receptor antagonists, such as BIBP-3226 (Amgen), BIBO-3304 (Boehringer Ingelheim), BMS-192548 and AR-H040922 (Bristol-Myers Squibb), LY-357897 (Eli Lilly), 1229U91 and GW438014S (GlaxoSmithKline), JNJ-5207787 (Johnson & Johnson), Lu-AA-44608 (Lundbeck), MK-0557 (Merck NPY), NGD-95-1 (Neurgogen), NLX-E201 (Neurologix), CGP-71683 (Novartis), PD-160170 (Pfizer), SR-120819A, BIIE0246, and S.A.0204 (Sanofi Aventis), S-2367 (Shiongli), dihydropyridine and dihydropyridine derivatives that are NPY receptor antagonists, bicyclic compounds that are NPY receptor antagonists, carbazole NPY receptor antagonists, and tricyclic compounds that are NPY receptor antagonists. *See, e.g.*, WO 2006/133160 and U.S. 6,313,128 (incorporated herein by reference in its entirety including figures).

**[0095]** Exemplary fat selective pro-apoptotic peptides includes, but is not limited to, CKGGRAKDC peptide that homes to white fat vasculature. *See*, Kolonin M.G. *et al.*, Nat. Med. June 10(6):625-32 (2004).

**[0096]** The compounds of preferred embodiments can be prepared from readily available starting materials using the following general methods and procedures. It will be appreciated that where typical or preferred process conditions (i.e., reaction temperatures, times, mole ratios of reactants, solvents, pressures, etc.) are given, other process conditions can also be used unless otherwise stated. Optimum reaction conditions may vary with the particular reactants or solvent used, but such conditions can be determined by one skilled in the art by routine optimization procedures.

[0097] Additionally, as will be apparent to those skilled in the art, conventional protecting groups may be necessary to prevent certain functional groups from undergoing undesired reactions. Suitable protecting groups for various functional groups as well as suitable conditions for protecting and deprotecting particular functional groups are well known in the art. For example, numerous protecting groups are described in T. W. Greene and G. M. Wuts, *Protecting Groups in Organic Synthesis*, Third Edition, Wiley, New York, 1999, and references cited therein.

[0098] The starting materials and reagents for the reactions described herein are generally known compounds or can be prepared by known procedures or obvious modifications thereof. For example, many of the starting materials and reagents are available from commercial suppliers such as Aldrich Chemical Co. (Milwaukee, Wisconsin, USA), Bachem (Torrance, California, USA), Emka-Chem or Sigma (St. Louis, Missouri, USA). Others may be prepared by procedures, or obvious modifications thereof, described in standard reference texts such as Fieser and Fieser's *Reagents for Organic Synthesis*, Volumes 1-15 (John Wiley and Sons, 1991), Rodd's *Chemistry of Carbon Compounds*, Volumes 1-5 and Supplementals (Elsevier Science Publishers, 1989), *Organic Reactions*, Volumes 1-40 (John Wiley and Sons, 1991), March's *Advanced Organic Chemistry*, (John Wiley and Sons, 4<sup>th</sup> Edition), and Larock's *Comprehensive Organic Transformations* (VCH Publishers Inc., 1989).

[0099] The various starting materials, intermediates, and compounds of the preferred embodiments may be isolated and purified where appropriate using conventional techniques such as precipitation, filtration, crystallization, evaporation, distillation, and chromatography. Characterization of these compounds may be performed using conventional methods such as by melting point, mass spectrum, nuclear magnetic resonance, and various other spectroscopic analyses.

### Examples

[0100] The various starting materials, intermediates, and compounds of the preferred embodiments may be isolated and purified where appropriate using conventional techniques such as precipitation, filtration, crystallization, evaporation, distillation, and chromatography. Characterization of these compounds may be performed using conventional methods such as by melting point, mass spectrum, nuclear magnetic resonance, and various other spectroscopic analyses.



[0101] Exemplary embodiments of steps for performing the synthesis of products in Synthetic Route #1, Scheme 1B is described in greater detail *infra*. Table 1 describes abbreviations used to express various compounds/moieties/apparatus/procedure/property in the exemplary reaction schemes and synthetic routes described in the following examples and throughout the specification.

**Table 1**

AcOH	Acetic acid
CAN	Acetonitrile
Ac <sub>2</sub> O	Acetic anhydride
AcCl	Acetyl chloride
NH <sub>4</sub> Cl	Ammonium chloride
CHCl <sub>3</sub>	Chloroform
CrO <sub>3</sub>	Chromium trioxide
DCA	Deoxycholic acid
DCM (CH <sub>2</sub> Cl <sub>2</sub> )	Dichloromethane
DMF	<i>N, N</i> -Dimethylformamide
DMSO	Dimethyl sulfoxide
EtOAc	Ethyl acetate
EtAlCl <sub>2</sub>	Ethyl aluminum dichloride
Hz	Hertz
HPLC	High pressure liquid chromatography
HCl	Hydrochloric acid
LAH	Lithium aluminum hydride
LiOH	Lithium hydroxide
MgSO <sub>4</sub>	Magnesium sulfate
MHz	Megahertz
MeOH	Methanol



mmol	millimole
mL	milliliter
mol	mole
Obs	Observed
HClO <sub>4</sub>	Perchloric acid
PtO <sub>2</sub>	Platinum oxide
KBr	Potassium bromide
<i>K</i> -O <sup>t</sup> Bu	Potassium <i>tert</i> -butoxide
PCC	Pyridinium chlorochromate
Rep	Reported
NaOH	Sodium hydroxide
THF	Tetrahydrofuran
SOCl <sub>2</sub>	Thionyl chloride
TEA	Triethylamine
TLC	Thin layer chromatography
Wt	Weight

**[0102] General:** Manipulations of oxygen- and moisture-sensitive materials are conducted with two-necked flame dried flasks under an argon atmosphere. Column chromatography is performed using SE-Make silica gel (60-120 Mesh), Spectrochem silica gel (230-400 Mesh) or aluminium oxide 90-neutral. (SD-Fine Chem. Ltd., India). Analytical thin layer chromatography (TLC) was performed on Merck Kieselgel 60 F<sub>254</sub> (0.25 mm) plates (Merck & Co., Whitehouse Station, NJ). Visualization of spots was detected either by UV light (254 nm) lamp or by charring with a solution of sulfuric acid (5%) and *p*-anisaldehyde (3%) in ethanol.

**[0103] Apparatus:** Analysis of the compounds and products of the reaction schemes and synthetic routes described herein may be performed on the apparatus and equipment described *infra*.

### Nuclear Magnetic Resonance (NMR)

[0104] Proton and carbon nuclear magnetic resonance spectra ( $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR) are recorded on a Varian Mercury-Gemini 200 ( $^1\text{H}$  NMR, 200 MHz;  $^{13}\text{C}$  NMR, 50 MHz) or a Varian Mercury-Inova 500 ( $^1\text{H}$  NMR, 500 MHz;  $^{13}\text{C}$  NMR, 125 MHz) (Varian, Inc., Palo Alto, CA) spectrometer with solvent resonances as the internal standards ( $^1\text{H}$  NMR,  $\text{CHCl}_3$  at 7.26 ppm or DMSO at 2.5 ppm and DMSO- $\text{H}_2\text{O}$  at 3.33 ppm;  $^{13}\text{C}$  NMR,  $\text{CDCl}_3$  at 77.0 ppm or DMSO at 39.5 ppm).  $^1\text{H}$  NMR data are reported as follows: chemical shift ( $\delta$ , ppm), multiplicity (s = singlet, d = doublet, t = triplet, q = quartet, br = broad, m = multiplet), coupling constants (Hz), and integration.

### Infrared Spectroscopy

[0105] Infrared spectra (FT-IR) are run on a JASCO-460<sup>+</sup> model (Jasco, Inc., Easton, MD). Mass spectra are obtained with a Perkin Elmer, API-2000 spectrometer (Perkin Elmer, Inc., Waltham, MA) using  $\text{ES}^+$  mode.

### Melting Point

[0106] Melting points were determined using a LAB-INDIA melting point measuring apparatus (Labindia Instruments Pvt. Ltd., India) and are uncorrected.

### High Pressure Liquid Chromatography

[0107] HPLC chromatograms were recorded using a SHIMADZU-2010 model with a PDA detector (Shimadzu Corp., Japan).

### Optical Activity

[0108] Specific optical rotations ( $[\alpha]_D$ ) are determined employing a JASCO-1020 at 589 nm (Jasco, Inc., Easton, MD) and are uncorrected.

[0109] **Chemicals:** Unless otherwise noted, commercially available reagents are used without purification. Diethyl ether and THF are distilled from sodium/benzophenone ketyl. Anhydrous DMF, DCM, pentane and hexane are obtained by distillation from  $\text{CaH}_2$ .

### Example 1: Preparation of Androstane-3,11,17-trione (1.13):

[0110] 10% of Pd/C (2.5 g, 5 wt %) is added to a solution of hydrocortisone (Compound 1.12) (50.0 g, 138.12 mmol) in DMF (250 mL). The resulting slurry is hydrogenated in a Parr apparatus (50 psi) for 12 h. Upon complete disappearance of

starting material, as evidenced by TLC, the crude reaction mixture is filtered through a small plug of Celite, and the solvent is removed under vacuum. Crude product (48.0 g) is obtained as a colorless solid.

[0111] NaBH<sub>4</sub> (2.1 g, 55.3 mmol) is added to a solution of the above crude product (48.0g, 131.86 mmol) in EtOH (500 mL) and CH<sub>2</sub>Cl<sub>2</sub> (500 mL). After 1 hr, acetone (50 mL) and water (150 mL) are added, followed by NaIO<sub>4</sub> (70.5 g, 329.6 mmol). The mixture is stirred at room temperature overnight.

[0112] Distilled water (500 mL) is added and the mixture is extracted with ethyl acetate (3 × 250 mL). The ethyl acetate layer is flushed through a silica-gel plug and the solvent is evaporated to yield 38 g as a colorless solid. The crude product is oxidized further without purification.

[0113] PCC (40.4 g, 187.5 mmol) is added to a solution of the above crude product in CH<sub>2</sub>Cl<sub>2</sub> (400 mL) in 3 equal portions over 30 minutes. The resulting reaction mixture is stirred at room temperature for about 3-4 h. Upon completion of the reaction, as monitored by TLC, the crude reaction mixture is filtered sequentially through pads of Celite and silica gel and the crude material is purified by column chromatography [59(W)×700(L) mm, 60-120 Mesh silica, 150 g], eluting with ethyl acetate/hexane (3:10) [50 mL fractions, 10 mL/min elution, monitored by TLC with *p*-anisaldehyde charring; R<sub>f</sub> for Compound **1.13** = 0.37 and R<sub>f</sub> for Compound **1.12** = 0.05 in EtOAc/Hexane (1 : 1)] to provide the diastereomeric Compound **1.13** (33.0 g, 79% yield) as a colorless solid.

[0114] The obtained crude material was purified by preparative HPLC using a Phenomenex Lunov C18 column (250 × 30.0 mm, 10μ) and isocratic elution with CH<sub>3</sub>CN : H<sub>2</sub>O (12 : 13) with a 25 mL/min flow rate in 15 mL fractions. The preparative HPLC is only used for purification, but not for analysis. Table 2 describes the measured properties of the product.

Table 2

<sup>1</sup> H NMR (500 MHz, CDCl <sub>3</sub> )	δ = 2.76 (dt, <i>J</i> = 4.0, 15.0 Hz, 1H), 2.62-2.35 (m, 5H), 2.33-2.24 (m, 1H), 2.23-2.05 (m, 4H), 2.02-1.88 (m, 3H), 1.81 (bd, <i>J</i> = 14.0 Hz, 2H), 1.72-1.61 (m, 1H), 1.57-1.48 (m, 1H), 1.47-1.32 (m, 2H), 1.26 (s, 3H), 0.86 (s, 3H)
<sup>13</sup> C NMR (125 MHz, CDCl <sub>3</sub> )	δ = 216.9, 211.8, 208.4, 52.3, 50.3, 50.2, 50.0, 44.5, 41.9, 37.1, 36.0, 35.9, 35.8, 34.3, 25.6, 25.0, 22.2, 21.3, 14.5



Mass (m/z)	303.2 [M <sup>+</sup> + 1], 320.1 [M <sup>+</sup> + 18]
IR (KBr)	3443, 2916, 1729, 1705, 1466, 1379, 1044 cm <sup>-1</sup>
m.p.	128.9-131 °C (from CH <sub>2</sub> Cl <sub>2</sub> /Hexane) (observed); 128-131 °C (Rep. E.Caspi <i>J. Org. Chem.</i> <b>1959</b> , 24, 669)
[α] <sub>D</sub>	+139 (c = 1 in CHCl <sub>3</sub> ).
HPLC purity	98.6%, ret. time = 16.61, (Hypersil BDS C18; 250 × 4.6 mm, 5u), ACN: 5 mM TEA pH-2.5 with HClO <sub>4</sub> (Gradient), absorbance at 205 nm

**Example 2: 3β-Hydroxy-androstane-11,17-dione (1.14):**

[0115] *K*-selectride<sup>®</sup> (98.39 mL, 98.01 mmol, 1M solution in THF) is added to a solution of Compound **1.13** (33.0 g, 109.27 mmol) in THF (330 mL) over 15 minutes under an inert atmosphere at −78 °C and is stirred for about 3-4 h at −78 °C. The reaction mixture is quenched with aqueous NaOH solution (2M, 70 mL). The crude reaction mixture is diluted with ethyl acetate (500 mL) and the organic layer is washed with water (3 × 75 mL), saturated brine solution (100 mL) and dried over MgSO<sub>4</sub> (75 g). The solvent is removed under vacuum to afford 33 g of crude material. The crude product is subjected to acetylation without purification.

***Purification of Crude Material***

[0116] The crude material is purified by column chromatography [29(W)×600(L) mm, 230-400 Mesh silica, 200 g], eluting with ethyl acetate/hexane (1 : 4) [25 mL fractions, 5 mL/min elution, monitored by TLC with *p*-anisaldehyde charring; R<sub>f</sub> for Compound **1.14** = 0.3 and R<sub>f</sub> for Compound **1.13** = 0.37 in EtOAc/Hexane (1 : 1)] to afford Compound **1.14**. Table 3 describes the measured properties of the product.

Table 3

<sup>1</sup> H NMR (500 MHz, CDCl <sub>3</sub> )	δ = 4.08 (s, 1H), 2.53 (q, <i>J</i> = 9.0 Hz, 1H), 2.42 (d, <i>J</i> = 13.0 Hz, 1H), 2.34-2.21 (m, 3H), 2.11-2.04 (m, 1H), 1.98-1.91 (m, 3H), 1.88-1.59 (m, 6H), 1.57-1.26 (m, 6H), 1.21 (s, 3H), 0.82 (s, 3H)
<sup>13</sup> C NMR (125 MHz, CDCl <sub>3</sub> )	δ = 217.4, 209.1, 66.3, 51.6, 50.6, 50.5, 37.1, 36.2, 35.9, 34.8, 33.5, 28.8, 28.5, 25.7, 25.6, 23.6, 21.5, 14.5
Mass (m/z)	305.0 [M <sup>+</sup> + 1], 322.0 [M <sup>+</sup> + 18]
IR (KBr)	3519, 2928, 1735, 1697, 1454, 1379 cm <sup>-1</sup>
m.p.	176.6-180.5 °C
[α] <sub>D</sub>	+125 (c = 1 in CHCl <sub>3</sub> )

**Example 3: 3 $\beta$ -Hydroxyandrosterone-11,17-dione acetate (1.15):**

[0117] Acetic anhydride (16.6 g, 162.8 mmol) is added to a solution of Compound **1.14** (33.0 g, 108.55 mmol) in pyridine (150 mL) at 0 °C under an inert atmosphere. The resulting reaction mixture is stirred overnight at ambient temperature. Upon completion of the reaction, as evidenced by TLC, pyridine and remaining acetic anhydride are removed under vacuum. The crude residue is diluted with ethyl acetate (500 mL) and washed with water (3  $\times$  150 mL), saturated brine solution (100 mL) and dried over MgSO<sub>4</sub> (75 g). The solvent is evaporated under vacuum and the crude material is purified by column chromatography [59(W) $\times$ 800(L) mm, 60-120 Mesh silica, 150 g], eluting with ethyl acetate/hexane (1:10) [25 mL fractions, 10 mL/min elution, monitored by TLC with *p*-anisaldehyde charring; R<sub>f</sub> for Compound **1.15** = 0.38 and R<sub>f</sub> for Compound **1.14** = 0.1 in EtOAc/Hexane (3 : 7)] to afford Compound **1.15** (19.0 g, 66.4 % yield) as a colorless solid. Table 4 describes the measured properties of the product.

Table 4

<sup>1</sup> H NMR (500 MHz, CDCl <sub>3</sub> ):	$\delta$ = 5.03 (s, 1H), 2.53 (dd, <i>J</i> = 9.5, 19.0 Hz, 1H), 2.42 (d, <i>J</i> = 10.0 Hz, 1H), 2.36-2.31 (m, 3H), 2.25 (dd, <i>J</i> = 9.5, 19.0 Hz, 1H), 2.10-2.06 (m, 1H), 2.04 (s, 3H), 1.96-1.91 (m, 3H), 1.81-1.69 (m, 2H), 1.63-1.57 (m, 3H), 1.50 (dd, <i>J</i> = 3.0, 14.5 Hz, 1H), 1.36 (d, <i>J</i> = 9.5 Hz, 3H), 1.27-1.22 (m, 1H), 1.20 (s, 3H), 0.82 (s, 3H)
<sup>13</sup> C NMR (125 MHz, CDCl <sub>3</sub> )	$\delta$ = 217.2, 208.9, 170.4, 69.7, 51.5, 50.5, 50.4, 37.9, 36.1, 35.9, 34.5, 30.6, 29.6, 29.5, 25.5, 25.4, 25.3, 23.4, 21.4, 21.3, 14.5
Mass (m/z)	347.1 [M <sup>+</sup> + 1], 364.1 [M <sup>+</sup> + 18]
IR (KBr)	3455, 2927, 1737.6, 1720.2, 1707.7, 1259, 1244 cm <sup>-1</sup>
m.p.	156-158 °C
[ $\alpha$ ] <sub>D</sub>	116 (c = 1 in CHCl <sub>3</sub> )

**Example 4: (Z)-3 $\beta$ -Hydroxy-5 $\beta$ -preg-17(20)-ene-11-one acetate (1.16):**

[0118] Potassium *tert*-butoxide (159.28 mL, 159.2 mmol, 1M solution in THF) is added to a solution of ethyltriphenylphosphonium bromide (61.16 g, 164.8 mmol) in THF (150 mL) is added drop wise over 1 h under an inert atmosphere at -5 °C. The resulting dark pink colored reaction mixture is warmed to 10-15 °C and stirred for an additional 1 h at the same temperature. A solution of Compound **55** (19.0 g, 54.9 mmol) in THF (50 mL) is introduced slowly to the above Wittig ylide suspension at -5 °C. The solution is stirred for an additional 10-20 minutes and the reaction mixture is allowed to warm to ambient temperature slowly. Stirring is continued for about 3-4 h. Upon complete disappearance of starting material, as evidenced by TLC, the reaction mixture is quenched with saturated



aqueous  $\text{NH}_4\text{Cl}$  solution (75 mL). The aqueous layer is extracted with EtOAc ( $2 \times 150$  mL) and the combined organic extracts are washed with saturated brine solution (100 mL) and dried over  $\text{MgSO}_4$  (75 g). The solvent is removed under vacuum and the crude material is purified by column chromatography [49(W) $\times$ 600(L) mm, 60-120 Mesh silica, 300 g] eluting with ethyl acetate/hexane (1 : 20) [25 mL fractions, 10 mL/min elution, monitored by TLC with *p*-anisaldehyde charring;  $R_f$  for Compound **1.16** = 0.54 and  $R_f$  for Compound **1.15** = 0.06 in EtOAc/Hexane (1 : 6)] to afford Compound **1.16** (15.5 g, 78.8% yield) as a thick colorless liquid, which solidified slowly after 1-2 days at 0 °C. Table 5 describes the measured properties of the product.

Table 5

$^1\text{H}$ NMR (500 MHz, $\text{CDCl}_3$ )	$\delta$ = 5.20-5.15 (m, 1H), 5.03 (s, 1H), 2.86 (d, $J$ = 10.0 Hz, 1H), 2.60 (d, $J$ = 10.0 Hz, 1H), 2.46-2.28 (m, 5H), 2.01 (s, 3H), 1.94-1.62 (m, 5H), 1.60-1.52 (m, 6H), 1.48-1.45 (m, 1H), 1.41-1.36 (m, 4H), 1.20 (s, 3H), 0.82 (s, 3H)
$^{13}\text{C}$ NMR (125 MHz, $\text{CDCl}_3$ )	$\delta$ = 210.9, 170.5, 147.2, 114.7, 70.0, 56.1, 55.6, 51.5, 47.4, 37.9, 35.9, 34.4, 31.6, 30.7, 29.8, 26.4, 25.9, 25.5, 24.0, 23.5, 21.3, 17.9, 12.8
Mass (m/z)	359.2 [ $\text{M}^+ + 1$ ], 376.2 [ $\text{M}^+ + 18$ ]
IR ( $\text{CHCl}_3$ )	3421, 2928, 1734, 1704, 1377, 1243 $\text{cm}^{-1}$
m.p.	88.5-91.2 °C
$[\alpha]_D$	+30 ( $c$ = 1 in $\text{CHCl}_3$ )

**Example 5: Methyl (*E*)-3 $\beta$ -hydroxy-5 $\beta$ -chola-16(17),22(23)-diene-24-oate acetate (1.17):**

[0119] Methyl propiolate (9.68 g, 114.95 mmol) is added to a solution of Compound **1.16** (16.5 g, 46 mmol) in  $\text{CH}_2\text{Cl}_2$  (220 mL) 0 °C. The reaction mixture is warmed to ambient temperature and stirred for 1 h under an inert atmosphere. Ethyl aluminum dichloride (17.5 g, 137.8 mmol) is introduced to the above mixture at 0 °C drop wise and the resulting reaction mass is again warmed to ambient temperature and stirred overnight. Upon completion of the reaction, as evidenced by TLC, the crude reaction mixture is quenched with ice-water (100 mL) and the aqueous layer is extracted with EtOAc ( $3 \times 150$  mL). The combined organic layer is washed with saturated brine solution (100 mL) and dried over  $\text{MgSO}_4$  (50 g). The solvent is removed under vacuum and the crude material is purified by column chromatography [49(W) $\times$ 600(L) mm, 60-120 Mesh silica, 300 g] eluting with ethyl acetate/hexane (1 : 7) [15 mL fractions, 10 mL/min elution, monitored by TLC and detected with either by UV light (254 nm) lamp or *p*-anisaldehyde



charring;  $R_f$  for Compound **1.17** = 0.36 and  $R_f$  for Compound **1.16** = 0.54 in EtOAc/Hexane (1 : 6)] to afford Compound **1.17** (16 g, 79% yield) as a colorless semi solid. Table 6 describes the measured properties of the product.

Table 6

$^1\text{H}$ NMR (500 MHz, $\text{CDCl}_3$ )	$\delta$ = 6.89 (dd, $J$ = 8.0, 16 Hz, 1H), 5.81 (d, $J$ = 15 Hz, 1H), 5.48 (s, 1H), 5.03 (s, 1H), 3.73 (s, 3H), 2.95 (t, $J$ = 6.5 Hz, 1H), 2.45-2.36 (m, 2H), 2.30-2.17 (m, 2H), 2.04 (s, 3H), 2.00-1.79 (m, 5H), 1.58 (s, 3H), 1.49-1.18 (m, 9H), 1.16 (s, 3H), 0.70 (s, 3H)
$^{13}\text{C}$ NMR (125 MHz, $\text{CDCl}_3$ )	$\delta$ = 209.7, 170.1, 166.6, 154.4, 152.3, 124.5, 119.1, 69.7, 56.3, 53.8, 52.3, 51.1, 49.8, 37.9, 35.7, 35.0, 34.4, 30.5, 30.4, 29.6, 26.2, 25.6, 25.2, 23.3, 21.1, 19.2, 17.3
Mass (m/z)	443.0 [ $\text{M}^+$ + 1], 460.1 [ $\text{M}^+$ + 18]
IR ( $\text{CHCl}_3$ )	3438, 2930, 1729, 1706, 1653, 1448, 1435, 1243, 1022 $\text{cm}^{-1}$
$[\alpha]_D$	+59 ( $c$ = 1 in $\text{CHCl}_3$ )
HPLC purity	94.4%; ret. time = 28.86, (Zorbax SB, C18; 250 $\times$ 4.6 mm, 5u), ACN: 5 mM TEA pH-2.5 with $\text{HClO}_4$ (Gradient); absorbance at 205 nm

**Example 6: Methyl 3 $\beta$ -hydroxy-5 $\beta$ -cholan-11-one-24-oate acetate (1.18):**

[0120] 10% Pd/C (2.9 g, 20 wt%) is added to a solution of Compound **1.17** (14.5 g, 32.8 mmol) in EtOAc (150 mL). The resulting slurry is hydrogenated in a Parr apparatus (50 psi) for 12 h. Upon complete disappearance of starting material, as evidenced by TLC [ $R_f$  for Compound **1.18** = 0.43 and  $R_f$  for Compound **1.17** = 0.43 in EtOAc/Hexane (1 : 3); however only Compound **1.17** is UV active from the conjugated ester chromophore], the crude reaction mixture is filtered through a small plug of Celite and the solvent is removed under vacuum to afford Compound **1.18** (14 g, 95.7% yield) as a colorless solid. Table 7 describes the measured properties of the product.

Table 7

$^1\text{H}$ NMR (500 MHz, $\text{CDCl}_3$ )	$\delta$ = 5.03 (s, 1H), 3.73 (s, 3H), 2.56 (d, $J$ = 10 Hz, 1H), 2.38-2.19 (m, 5H), 2.04 (s, 3H), 1.86-1.13 (m, 20H), 1.12 (s, 3H), 0.86 (s, 3H), 0.62 (s, 3H)
$^{13}\text{C}$ NMR (125 MHz, $\text{CDCl}_3$ )	$\delta$ = 211.3, 174.3, 170.5, 70.0, 58.3, 55.7, 55.0, 51.4, 50.8, 46.8, 37.9, 36.7, 35.0, 34.3, 30.9, 30.7, 30.7, 29.7, 28.3, 26.6, 25.9, 25.5, 23.6, 23.5, 21.4, 17.9, 12.7
Mass (m/z)	447.1 [ $\text{M}^+$ + 1], 464.1 [ $\text{M}^+$ + 18]
IR (KBr)	3449, 2927, 1734, 1704, 1381, 1262, 1243 $\text{cm}^{-1}$
m.p.	174.2-175.7 $^{\circ}\text{C}$ (From $\text{CH}_2\text{Cl}_2$ /Hexane) (Observed); 174.8-176.2 $^{\circ}\text{C}$ (Reported)

$[\alpha]_D$	+39 (c = 1 in CHCl <sub>3</sub> )
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**Example 7: Methyl 3 $\beta$ -hydroxy-5 $\beta$ -chol-9(11)-ene-24-oate acetate (1.19):**

**[0121]** PtO<sub>2</sub> (5.0 g, 100 wt%) is added to a solution of **1.18** (5.0 g, 11.2 mmol) in EtOAc (75 mL) in the presence of catalytic amount of AcOH (2.0 mL). The resulting slurry is hydrogenated in a Parr apparatus (70 psi) for about 14-16 h. Upon completion of the reaction, the crude mixture is filtered through a small plug of Celite and the solvent is removed under vacuum. The crude product is used for the elimination reaction without further purification.

**[0122]** SOCl<sub>2</sub> (1.98 g, 16.78 mmol) is introduced to a solution of the above crude material in pyridine (100 mL) drop wise at 0 °C. The resulting reaction mixture is warmed to ambient temperature and stirred for about 1 h. Upon completion of the reaction, as evidenced by TLC, pyridine is removed under vacuum. The crude residue is diluted with ethyl acetate (100 mL) and washed with water (2 × 50 mL), saturated brine solution (100 mL) and dried over MgSO<sub>4</sub> (40 g). The solvent is evaporated under vacuum and the crude material is purified by column chromatography [49(W)×600(L) mm, 60-120 Mesh silica, 120 g] eluting with ethyl acetate/hexane (1:10) [10 mL fractions, 5 mL/min elution, monitored by TLC with *p*-anisaldehyde charring; R<sub>f</sub> for Compound **1.19** = 0.51 and R<sub>f</sub> for Compound **1.18** = 0.22 in EtOAc/Hexane (1 : 6)] to afford Compound **1.19** (4.1 g, 85.4% yield) as a colorless solid. Table 8 describes the measured properties of the product.

Table 8

<sup>1</sup> H NMR (500 MHz, CDCl <sub>3</sub> )	δ = 5.33 (s, 1H), 5.03 (s, 1H), 3.66 (s, 3H), 2.37-2.32 (m, 1H), 2.46-2.21 (m, 1H), 2.11-2.04 (m, 1H), 2.03 (s, 3H), 1.99-1.10 (m, 22H), 1.07 (s, 3H), 0.92 (d, <i>J</i> = 7.0 Hz, 3H), 0.58 (s, 3H)
<sup>13</sup> C NMR (125 MHz, CDCl <sub>3</sub> )	δ = 174.6, 170.6, 140.1, 118.9, 71.2, 56.1, 53.3, 51.3, 42.0, 40.9, 38.9, 37.3, 36.4, 35.2, 31.9, 31.4, 31.0, 30.9, 30.1, 28.2, 26.9, 26.2, 26.1, 25.3, 21.4, 17.9, 11.6
Mass (m/z)	448.2 [M <sup>+</sup> + 18]
IR (KBr)	3447, 2935, 1735, 1379, 1261, 1245 cm <sup>-1</sup>
m.p.	188.6-191.2 °C (From CH <sub>2</sub> Cl <sub>2</sub> /Hexane) (Observed); 174-175 °C (Reported)
$[\alpha]_D$	+ 37 (c = 1 in CHCl <sub>3</sub> )



**Example 8: Methyl 3 $\beta$ -hydroxy-5 $\beta$ -chol-9(11)-ene-12-one-24-oate acetate (1.20):**

[0123] CrO<sub>3</sub> (8.0 g, 100 wt%, 80.0 mmol) is added to a solution of Compound **1.19** (8.0 g, 18.6 mmol) in AcOH (150 mL). The resulting reaction mixture is heated at 60 °C for about 24-36 h. Upon complete disappearance of the precursor, acetic acid is evaporated under vacuum, and the crude material is dissolved in diethyl ether (400 mL). The organic layer is washed with water (2 × 100 mL), saturated brine solution (100 mL) and dried over MgSO<sub>4</sub> (40 g). The solvent is removed under vacuum and the crude material is purified by column chromatography [49(W)×600(L) mm, 60-120 Mesh silica, 120 g] eluting with ethyl acetate/hexane (1 : 5) [10 mL fractions, 3 mL/min elution, monitored by TLC and detected with UV light (254 nm) lamp; R<sub>f</sub> for Compound **1.20** = 0.28 and R<sub>f</sub> for Compound **1.19** = 0.61 in EtOAc/Hexane (1 : 4)] to afford Compound **1.20** (5 g, 60.5% yield) as a colorless solid. Table 9 describes the measured properties of the product.

Table 9

<sup>1</sup> H NMR (500 MHz, CDCl <sub>3</sub> )	δ = 5.72 (s, 1H), 5.04 (s, 1H), 3.66 (s, 3H), 2.41-2.27 (m, 3H), 2.03 (s, 3H), 1.94-1.58 (m, 9H), 1.48-1.30 (m, 11H), 1.21 (s, 3H), 1.02 (d, <i>J</i> = 6.5 Hz, 3H), 0.91 (s, 3H)
<sup>13</sup> C NMR (500 MHz, CDCl <sub>3</sub> )	δ = 205.1, 174.5, 170.4, 164.2, 123.1, 70.3, 53.4, 53.0, 51.3, 47.2, 40.2, 37.7, 37.3, 35.2, 32.1, 31.4, 31.0, 30.6, 30.2, 27.3, 26.5, 25.9, 25.6, 24.1, 21.3, 19.4, 10.6
Mass (m/z)	445.0 [M <sup>+</sup> + 1], 462.1 [M <sup>+</sup> + 18]
IR	3447, 2927, 2361, 2339, 1736, 1678, 1367, 1250 cm <sup>-1</sup>
m.p.	185.8-188.1 °C (From CH <sub>2</sub> Cl <sub>2</sub> /Hexane)
[α] <sub>D</sub>	+62 (c = 1 in CHCl <sub>3</sub> )
HPLC purity	94.1%; ret. time = 23.89 (Hypersil BDS C18, 250 × 4.6 mm, 5u, CH <sub>3</sub> CN : 5 mM TEA, pH-2.5 with HClO <sub>4</sub> (Gradient); absorbance at 240 nm

**Example 9: Methyl 3 $\beta$ -hydroxy-5 $\beta$ -cholane-12-one-24-oate acetate (1.21):**

[0124] 10% Pd/C (30 mg, 10 wt%) is added to a solution of Compound **1.20** (300 mg, 0.675 mmol) in EtOAc (30 mL). The resulting slurry is hydrogenated in a Parr apparatus (50 psi) for about 16 h. Upon complete disappearance of starting material by TLC [R<sub>f</sub> for Compound **1.21** = 0.44 and R<sub>f</sub> for Compound **1.20** = 0.44 in EtOAc/Hexane (3 : 7); however only Compound **1.20** is UV active from its enone chromophore; additionally charring of Compound **1.20** is faint but Compound **1.21** is bright], the crude reaction mixture was filtered through a small plug of Celite and the solvent is removed under



vacuum to afford Compound **1.21** (270m g, 90% yield) as a colorless solid. Table 10 describes the measured properties of the product.

Table 10

<sup>1</sup> H NMR (500 MHz, CDCl <sub>3</sub> )	δ = 5.04 (s, 1H), 3.64 (s, 3H), 2.52-2.47 (m, 1H), 2.38-2.25 (m, 3H), 2.23-2.03 (m, 2H), 2.02(s, 3H), 1.99-1.71 (m, 8H), 1.49-1.11 (m, 12H), 1.05 (s, 3H), 1.01 (s, 3H), 0.85-0.84 (d, <i>J</i> = 7.0 Hz, 3H)
<sup>13</sup> C NMR (125 MHz, CDCl <sub>3</sub> )	δ = 214.7, 174.6, 170.5, 70.2, 58.7, 57.5, 51.4, 46.5, 43.7, 38.4, 36.9, 35.7, 35.6, 35.5, 31.3, 30.7, 30.5, 27.5, 26.4, 25.9, 24.8, 24.3, 23.2, 21.4, 18.6, 11.7
Mass (m/z)	447.0 [M <sup>+</sup> + 1], 464.0 [M <sup>+</sup> + 18]
IR (KBr)	3447, 2935, 1735, 1704, 1260, 1241 cm <sup>-1</sup>
m.p.	179.6-182.7 °C (From CH <sub>2</sub> Cl <sub>2</sub> /Hexane)
[α] <sub>D</sub>	+69 (c = 1 in CHCl <sub>3</sub> )

**Example 10: Methyl 5β-chola-3,12-dione-24-oate (1.22):**

[0125] NaOH (73 mg, 1.8 mmol) is added to a solution of Compound **2.1** (270 mg, 0.6 mmol) in MeOH (10 mL). The resulting reaction mixture is stirred for about 2 h at ambient temperature. Upon completion of the reaction, as evidenced by TLC, MeOH is removed under vacuum and the crude product is diluted with ethyl acetate (20 mL). The organic layer is washed with saturated brine solution (10 mL) and dried over MgSO<sub>4</sub> (5.0 g). The solvent is removed under vacuum and the crude material is used in the esterification reaction without purification.

[0126] SOCl<sub>2</sub> (0.1 mL, 1.35 mmol) is added drop-wise to a solution of the above crude material in MeOH (10 mL) 0 °C. The resulting reaction mixture is stirred at ambient temperature for about 1 h. Upon completion of the reaction, MeOH is removed under vacuum. The crude reaction mixture is diluted with EtOAc (30 mL), and the organic layer is washed with water (3 × 10 mL), saturated brine solution (15 mL) and dried over MgSO<sub>4</sub> (5 g). The solvent is evaporated under vacuum and the crude product is used for the oxidation reaction without purification.

[0127] PCC (1.0 g, 4.6 mmol) is introduced in 3 equal portions to a solution of the obtained ester in CH<sub>2</sub>Cl<sub>2</sub> (25 mL) over about 5 minutes. The resulting reaction mixture is stirred at ambient temperature for about 3-4 h. Upon completion of the reaction, as evidenced by TLC, the crude reaction mixture is filtered through a pad of Celite. The solvent is removed under vacuum and the crude material is purified by column chromatography [19(W)×400(L) mm, 60-120 Mesh silica, 45 g] eluting with ethyl

acetate/hexane (1 : 6) [10 mL fractions, 5 mL/min elution, monitored by TLC with *p*-anisaldehyde charring;  $R_f$  for Compound **1.22** = 0.57 and  $R_f$  for Compound **1.21** = 0.71 in EtOAc/Hexane (2 : 3)] to afford Compound **1.22** (170 mg, 70.8% yield) as a colorless solid. Table 11 describes the measured properties of the product.

Table 11

$^1\text{H}$ NMR (500 MHz, $\text{CDCl}_3$ )	$\delta$ = 3.66 (s, 3H), 2.64-2.57 (m, 2H), 2.55-2.19 (m, 4H), 2.17-2.02 (m, 3H), 1.99-1.21 (m, 17H), 1.11 (s, 3H), 1.05 (s, 3H), 0.86 (d, $J$ = 10.0 Hz, 3H)
$^{13}\text{C}$ NMR (125 MHz, $\text{CDCl}_3$ )	$\delta$ = 213.9, 211.8, 174.5, 58.5, 57.5, 51.4, 46.5, 44.2, 43.7, 42.1, 38.3, 36.9, 36.8, 35.6, 35.4, 31.3, 30.5, 27.4, 26.6, 25.4, 24.3, 22.1, 18.5, 11.7
Mass (m/z)	403.1 [ $\text{M}^+$ + 1], 420.2 [ $\text{M}^+$ + 18]
IR (KBr)	3457, 2925, 1737, 1708, 1216, 1176 $\text{cm}^{-1}$
m.p.	133.7-135.9 $^{\circ}\text{C}$ (From $\text{CH}_2\text{Cl}_2$ /Hexane) (Obs); 136.5-137.5 $^{\circ}\text{C}$ (Rep)
$[\alpha]_D$	+79 ( $c$ = 1 in $\text{CHCl}_3$ )

**Example 11: Methyl deoxycholate (1.22-Ester):**

[0128]  $\text{LiAlH}(\text{O}^t\text{Bu})$  (332mg, 1.3 mmol,) is introduced drop-wise to a solution of Compound **1.22** (150 mg, 0.37 mmol) in THF (10 mL) under an inert atmosphere at ambient temperature. After being stirred for about 4-5 hr, the reaction mixture is quenched with Aqueous HCl (2 mL, 1N) and the crude mixture is diluted with EtOAc (30 mL), washed with water (15 mL), saturated brine solution (10 mL) and dried over  $\text{MgSO}_4$  (3 g). The solvent is removed under vacuum, and the crude mass is purified by column chromatography [29(W)×500(L) mm, 230-400 Mesh silica, 50 g] eluting with  $\text{MeOH}/\text{CH}_2\text{Cl}_2$  (1 : 20) [5 mL fractions, 3 mL/min elution, monitored by TLC with *p*-anisaldehyde charring;  $R_f$  for Compound **1.22-ester** = 0.42 and  $R_f$  for Compound **1.22** = 0.85 in  $\text{MeOH}/\text{CH}_2\text{Cl}_2$  (1 : 9)] to afford methyl deoxycholate (Compound **1.22-ester**) (110 mg, 72.8% yield) as a colorless solid. Table 12 describes the measured properties of the product.



Table 12

<sup>1</sup> H NMR (500 MHz, CDCl <sub>3</sub> )	δ = 3.97 (s, 1H), 3.65 (s, 3H), 3.63-3.59 (m, 1H), 2.39-2.33 (m, 1H), 2.25-2.19 (m, 1H), 1.88-0.97 (m, 24H), 0.95 (d, <i>J</i> = 6.0 Hz, 3H), 0.90 (s, 3H), 0.67(s, 3H)
<sup>13</sup> C NMR (125 MHz, CDCl <sub>3</sub> )	δ = 174.7, 73.1, 71.7, 51.4, 48.2, 47.3, 46.5, 42.1, 36.4, 36.0, 35.2, 35.1, 34.1, 33.6, 31.1, 30.9, 30.4, 28.6, 27.4, 27.1, 26.1, 23.6, 23.1, 17.3, 12.7
Mass (m/z)	407.1 [M <sup>+</sup> + 1], 424.2 [M <sup>+</sup> + 18]
IR (KBr)	3419, 2937, 2864, 1740, 1724, 1448, 1377, 1043 cm <sup>-1</sup>
m.p.	58.0-60.0 °C (under re-crystallization)
[α] <sub>D</sub>	+36 (c = 1 in CHCl <sub>3</sub> )

**Example 11: Deoxycholic Acid:**

[0129] A solution of LiOH (23 mg, 0.55 mmol) in H<sub>2</sub>O (2.0 mL) is added to a solution of **1.22-ester** (110 mg, 0.27 mmol) in THF (4 mL). The resulting reaction mixture is stirred for about 2-3 h at ambient temperature. Upon disappearance of the ester by TLC [*R<sub>f</sub>* for Compound **DCA** = 0.35 and *R<sub>f</sub>* for Compound **1.22-ester** = 0.42 in MeOH/CH<sub>2</sub>Cl<sub>2</sub> (1 : 9)], the crude reaction mixture is diluted with ethyl acetate (10 mL) and triturated with saturated brine solution to obtain a clear separation of the organic layer. The organic layer was washed with saturated NH<sub>4</sub>Cl solution (10 mL), and dried over MgSO<sub>4</sub> (3.0 g). The solvent is removed under vacuum and any trace of water is removed by azeotropic with toluene (3 × 5 mL) to afford deoxycholic acid (**DCA**) (100 mg, 94.3% yield) as a colorless solid. Table 13 describes the measured properties of the product.

Table 13

<sup>1</sup> H NMR (500 MHz, DMSO)	δ = 3.77 (s, 1H), 3.38-3.33 (m, 1H), 2.20-2.15 (m, 1H), 2.08-2.02 (m, 1H), 1.92-0.90 (m, 24H), 0.84-0.83 (d, <i>J</i> = 5.0 Hz, 3H), 0.77 (s, 3H), 0.53 (s, 3H)
<sup>13</sup> C NMR (125 MHz, DMSO)	δ = 175.9, 71.0, 69.9, 47.4, 46.2, 45.9, 41.6, 36.3, 35.7, 35.1, 35.0, 33.8, 32.9, 31.2, 31.2, 30.2, 28.6, 27.2, 26.9, 26.1, 23.5, 23.0, 16.9, 12.4
Mass (m/z)	392 [M <sup>+</sup> , not detected], 410.2 [M <sup>+</sup> + 18]
IR	3445, 2931, 2867, 1694, 1636, 1043 cm <sup>-1</sup>
m.p.	173.2-175.5 °C (From THF/CH <sub>2</sub> Cl <sub>2</sub> ) (Observed); 174-176 °C (Reported, Alfa Aesar) and 171-174 °C (Reported, Aldrich)
[α] <sub>D</sub>	+50 [c = 1 in MeOH and CHCl <sub>3</sub> (1:1)]; +54° (c = 2 in ethanol) [Alfa Aesar]

[0130] The yield of products for the exemplary processes described in Examples 1 through 11 are described in Table 14.



Table 14: Overall Yield of Synthetic DCA Process

Compound	MP (°C) (observed)	MP (°C) Reported	% Yield	Notes
Hydrocortisone				
<b>1.13</b>	128.9-131.1	128.0-131.0	79.00	Hydrogenation, side chain cleavage, and PCC oxidation
<b>1.14</b>	176.6-180.5		Crude	K-Selectride <sup>®</sup> reduction
<b>1.15</b>	156.0-158.0		66.40	Acetylation
<b>1.16</b>	88.5-91.2		78.80	Wittig
<b>1.17</b>			79.00	Ene
<b>1.18</b>	174.2-175.7		95.80	Hydrogenation (Pd/C)
<b>1.19</b>	188.6-191.2	174.0-175.0	85.40	Thionyl chloride/pyridine dehydration (2-step yield)
<b>1.20</b>	185.8-188.1		60.50	Chromium trioxide allylic oxidation
<b>1.21</b>	179.6-182.7		90.00	Hydrogenation (Pd/C)
<b>1.22</b>	133.7-135.9	136.5-137.5	70.80	Hydrolysis, esterification, and oxidation
<b>1.22-ester</b>	58.0-60.0		72.80	LiAlH(O- <sup>t</sup> Bu) <sub>3</sub> reduction
<b>DCA</b>	173.2-175.5	174.0-176.0	94.33	Hydrolysis of ester
			7.00	Overall yield

**Example 12: Comparison of carbon content of synthetic DCA versus bovine DCA**

[0131] Synthetic DCA (three samples, prepared according to the methods disclosed above or according to WO 2008/157635 Examples 12-24 which is incorporated herein by reference in its entirety) and DCA purchased from SIGMA (Sigma Aldrich PO Box 14508 St. Louis, MO 63178), PIERCE (Pierce Protein Research Products PO Box 117, Rockford IL 61105 USA), and NZP (New Zealand Pharmaceuticals Limited, PO Box 1869, Palmerston North 4440, New Zealand) were analyzed for their percent fossil carbon and ppt (parts per trillion) <sup>14</sup>C carbon content. The percent fossil carbon provides a measure of the amount of carbon in the molecule originating from fossil fuel, which is expected to have very little <sup>14</sup>C due to decay of this isotope. Intermediates used in the synthesis of DCA are derived from fossil fuels. Conversely, DCA synthesized from animals is expected to have approximately 1 ppt <sup>14</sup>C and little to no fossil content. These expectations are borne out in the analyses as shown in Table 15. (Radiocarbon analyses were carried out according to the American Society for Testing Materials ASTM D6866

procedure (ASTM international, 100 Barr Harbor Drive, PO Box C700, West Conshohocken, PA 19428-2959).

Table 15: Comparison of carbon content of DCA isolated from bovines versus synthetic DCA

DCA	Fossil Carbon	ppt C <sup>14</sup>
Bovine DCA (SIGMA)	0%	1 ppt
Bovine DCA (PIERCE)	0%	1 ppt
Bovine DCA (NZP)	2%	1 ppt
Synthetic DCA (sample 1)	13%	0.87 ppt
Synthetic DCA (sample 2)	12%	0.88 ppt
Synthetic DCA (sample 3)	11%	0.89 ppt

[0132] Accordingly, in one embodiment provided is method for distinguishing synthetic DCA from naturally derived DCA based on their fossil-derived carbon content and/or <sup>14</sup>C content. In one aspect, provided is synthetic DCA having greater than 4% fossil-derived carbon. In another aspect, provided is synthetic DCA having less than 1 ppt <sup>14</sup>C or less than 0.9 ppt <sup>14</sup>C.

### Example 13: Synthesis of DCA having Fossil Carbons in the Backbone Ring System

[0133] The starting material for synthesizing DCA, such as cortisone and hydrocortisone, can be prepared completely or partially by using fossil derived compounds. For example, cortisone can be prepared according to Nemoto, et al., First Enantioselective Total Synthesis of (+)-Cortisone, *J. Org. Chem.*, 55(21):5625–5631(1990); Hu, et al., Application of Chiral Cationic Catalysts to Several Classical Syntheses of Racemic Natural Products Transforms Them into Highly Enantioselective Pathways, *J. Am. Chem. Soc.* 126(42):13708-13(2004); Woodward R. B., et al., The Total Synthesis of Cortisone, *Journal of the American Chemical Society* 73:4057-4057(1951) and references cited therein, which are hereby incorporated by reference in their entirety. Other methods for synthesizing the starting material are known in the art. When the so-prepared starting material is used as starting material for preparing DCA, the resulting DCA will have fossil carbons in the backbone, and thus can have a <sup>14</sup>C content of less than 0.87 ppt.



[0134] The following table shows DCA (which has a total of 24 carbon atoms) with exemplifying  $^{14}\text{C}$  content.

Fossil Carbon	$^{14}\text{C}$ content (ppt)	Fossil Carbon	$^{14}\text{C}$ content (ppt)
0	about 1	13	about 0.458
1	about 0.958	14	about 0.417
2	about 0.917	15	about 0.375
3	about 0.875	16	about 0.333
4	about 0.833	17	about 0.292
5	about 0.792	18	about 0.25
6	about 0.75	19	about 0.208
7	about 0.708	20	about 0.167
8	about 0.667	21	about 0.125
9	about 0.625	22	about 0.083
10	about 0.583	23	about 0.047
11	about 0.542	24	0 (undetectable)
12	about 0.50		

[0135] Using starting material, such as cortisone and hydrocortisone, prepared completely from fossil derived materials could provide DCA having an undetectable amount of  $^{14}\text{C}$  content.

[0136] It is completed that mixing appropriate amounts DCA prepared completely or partially from fossil derived starting materials with DCA derived from plant or plant based starting materials could provide DCA having any amount of  $^{14}\text{C}$  content from greater than 0 ppt to less than 1 ppt, for example, a  $^{14}\text{C}$  content of from greater than 0 ppt to less than 0.85 ppt. It is completed that mixing appropriate amounts cortisone prepared completely or partially from fossil derived starting materials with cortisone derived from plant or plant based starting materials could provide DCA having any amount of  $^{14}\text{C}$  content from greater than 0 ppt to less than 1 ppt, for example, a  $^{14}\text{C}$  content of from greater than 0 ppt to less than 0.85 ppt. It is also completed that mixing appropriate amounts hydrocortisone prepared completely or partially from fossil derived starting materials with hydrocortisone derived from plant or plant based starting materials could provide DCA having any amount of  $^{14}\text{C}$  content from greater than 0 ppt to less than 1 ppt, for example, a  $^{14}\text{C}$  content of from greater than 0 ppt to less than 0.86 ppt.

#### **Example 14: Preparation of DCA having a $^{14}\text{C}$ Content of Greater Than 1 ppt**

[0137] DCA having a  $^{14}\text{C}$  content of greater than 1 ppt can be prepared according to methods described above by using one or more starting material having an enriched  $^{14}\text{C}$



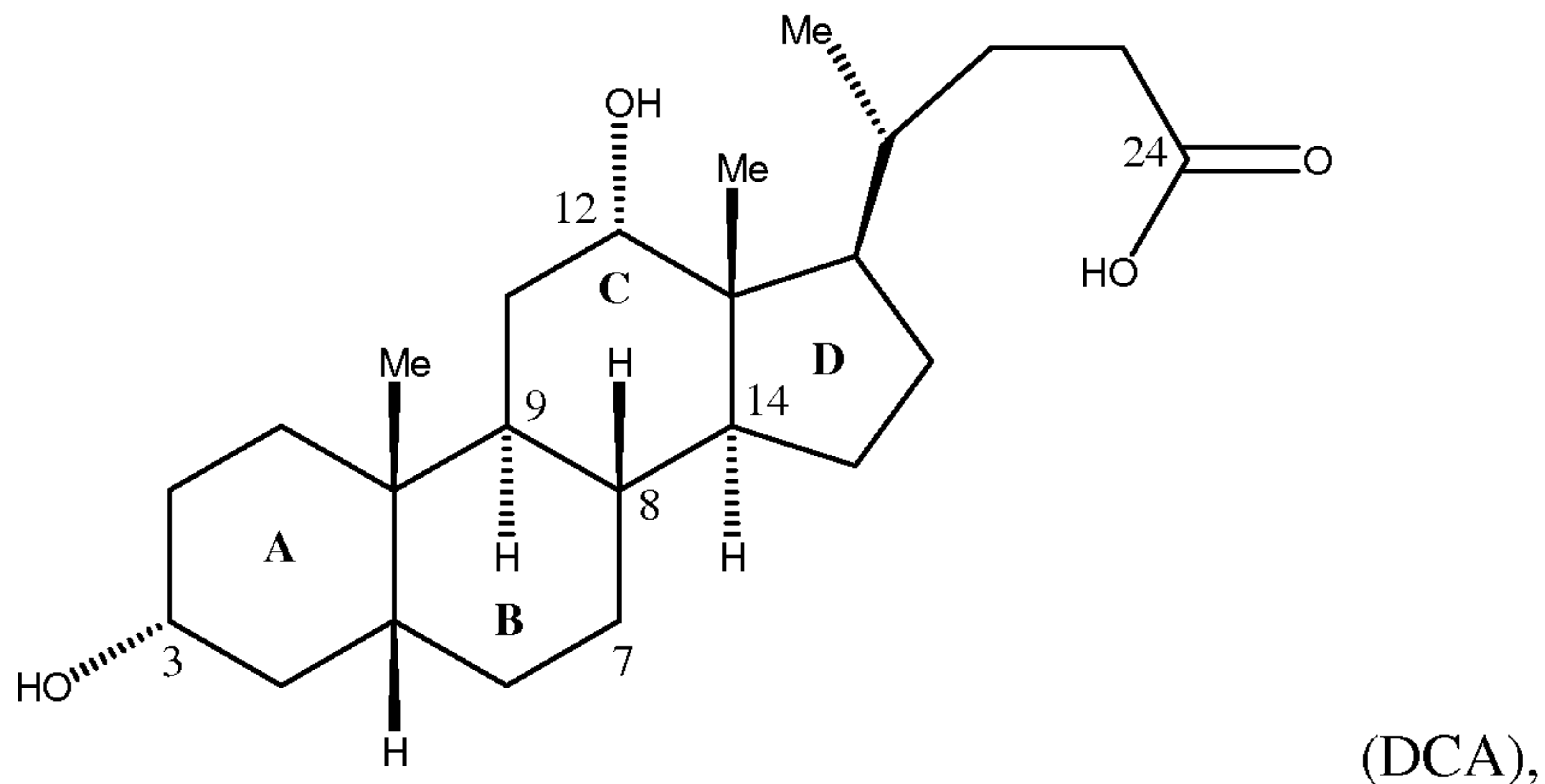
content. It is contemplated that the DCA with a greater  $^{14}\text{C}$  content is especially useful for monitoring its delivery, distribution, metabolism, clearance, and the like, when it is administered to a patient.

**Example 15: Preparation of Synthetic DCA having a  $^{14}\text{C}$  Content of 1 ppt**

[0138] Synthetic DCA having a  $^{14}\text{C}$  content of 1 ppt can be prepared by mixing an appropriate amount of DCA prepared partly or completely from fossil carbon starting materials with DCA having a  $^{14}\text{C}$  content of greater than 1 ppt prepared according to Example 14 above. The DCA so prepared would resemble the  $^{14}\text{C}$  content of naturally occurring DCA (e.g., the DCA isolated from animals) which is 1 ppt, and yet avoid the possibility of contamination by animal-borne pathogens.

**What is claimed is:**

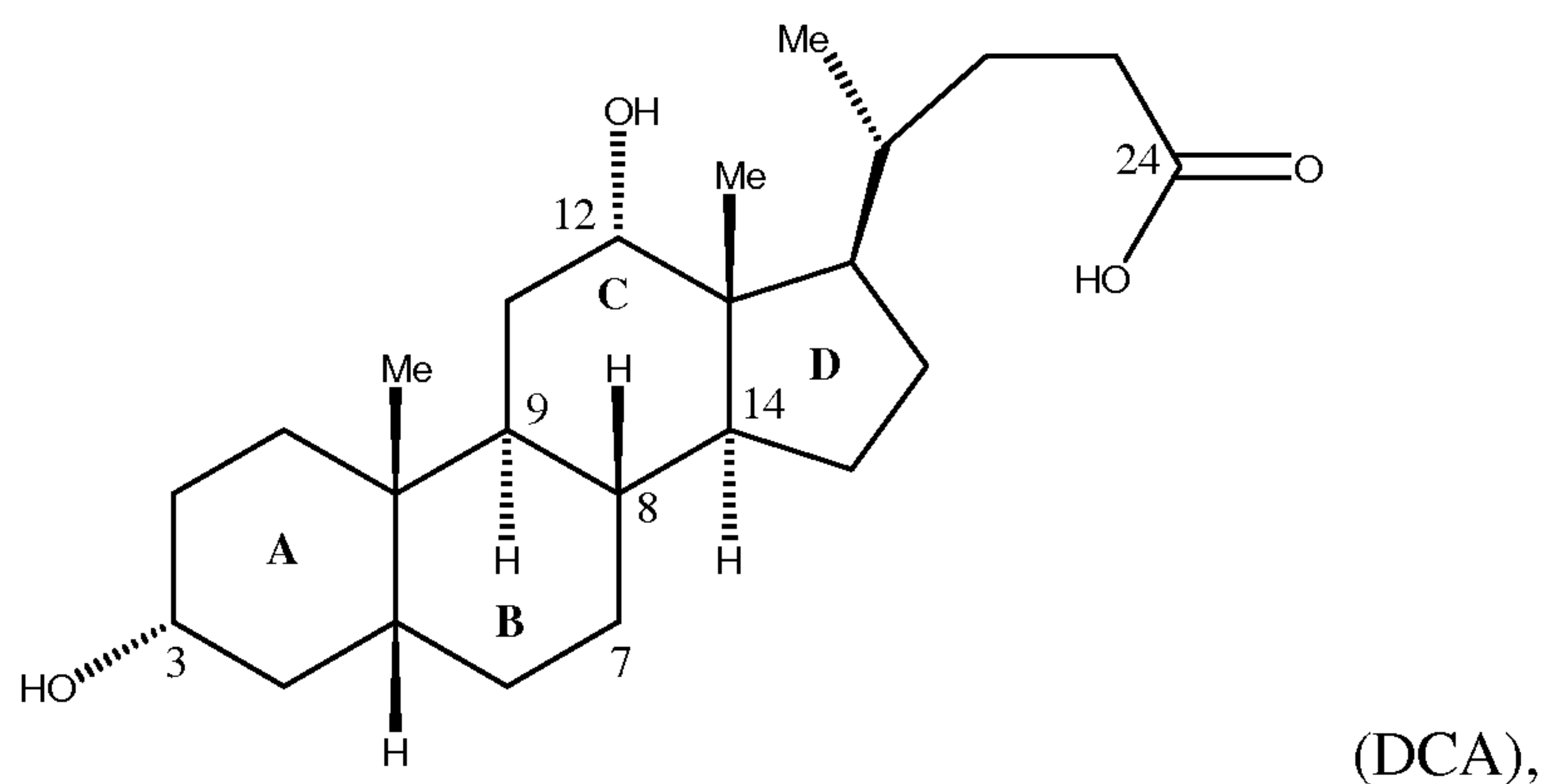
1. A compound that is deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



wherein said DCA has a  $^{14}\text{C}$  content of less than 1 ppt.

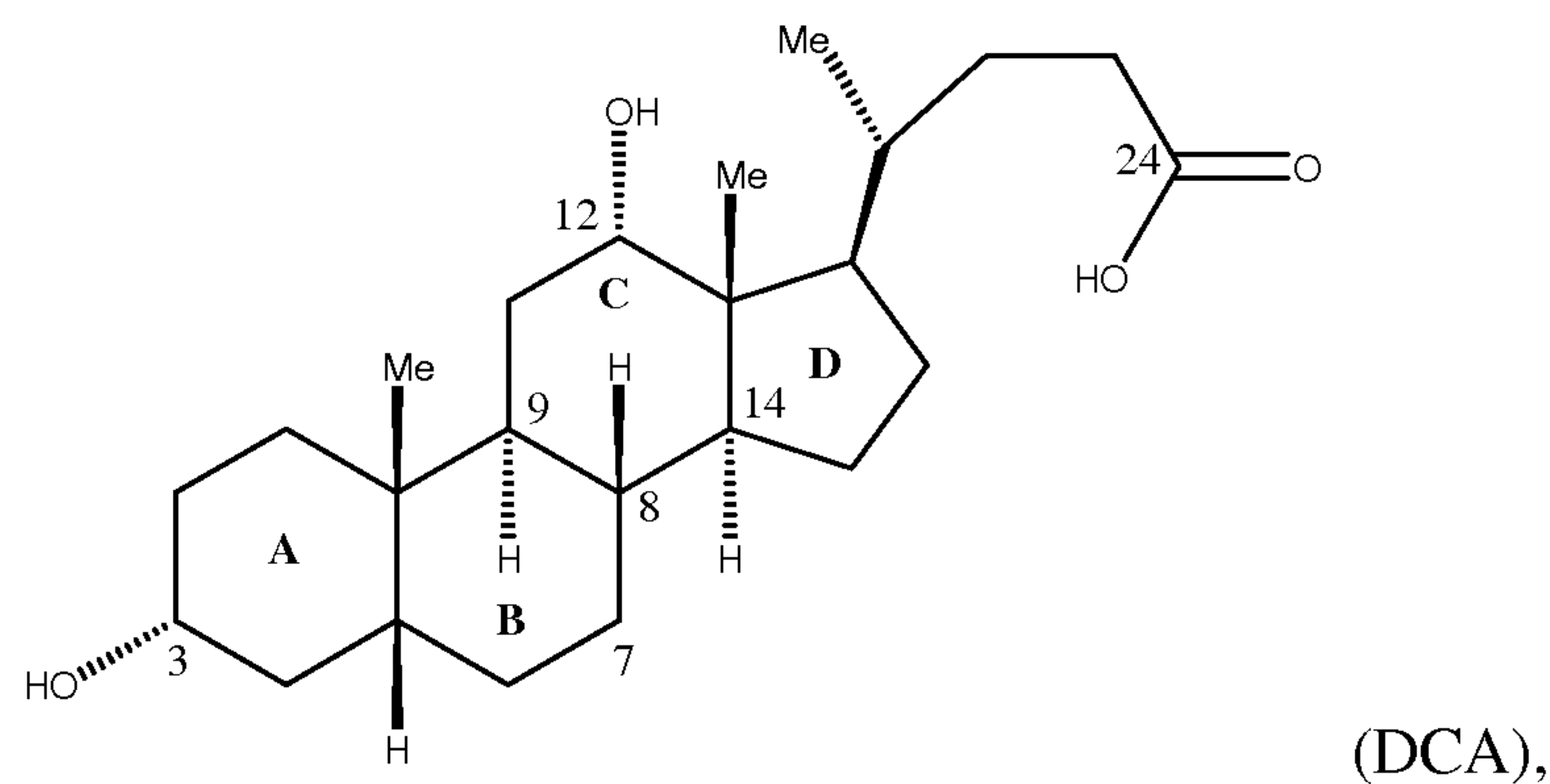
2. The compound of claim 1 wherein the DCA has a  $^{14}\text{C}$  content of less than 0.9 ppt.
3. The compound of claim 1 or 2, which is the salt of the DCA.
4. The compound of claim 3 which is a sodium salt of the DCA.
5. A composition comprising the compound of claim 1 and a pharmaceutically acceptable excipient.
6. The composition of claim 5 wherein the compound is the salt of the DCA.
7. The composition of claim 6 wherein the compound is the sodium salt of the DCA.

8. A compound that is deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



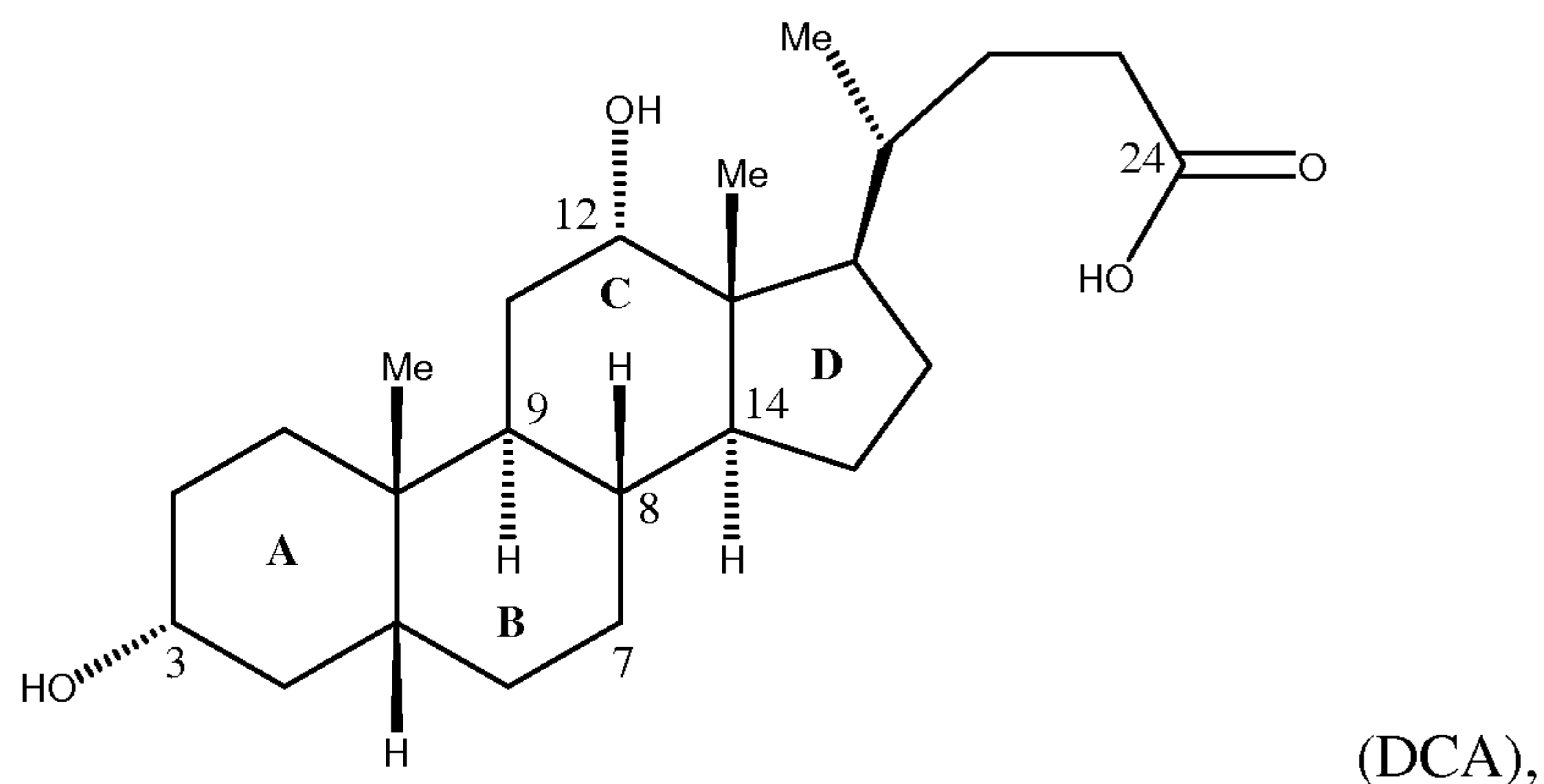
wherein said DCA has a  $^{14}\text{C}$  content of less than 0.86 ppt.

9. A compound that is deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



wherein said DCA has a  $^{14}\text{C}$  content of greater than 1 ppt.

10. A compound that is a synthetic deoxycholic acid (DCA) or a pharmaceutically acceptable salt thereof:



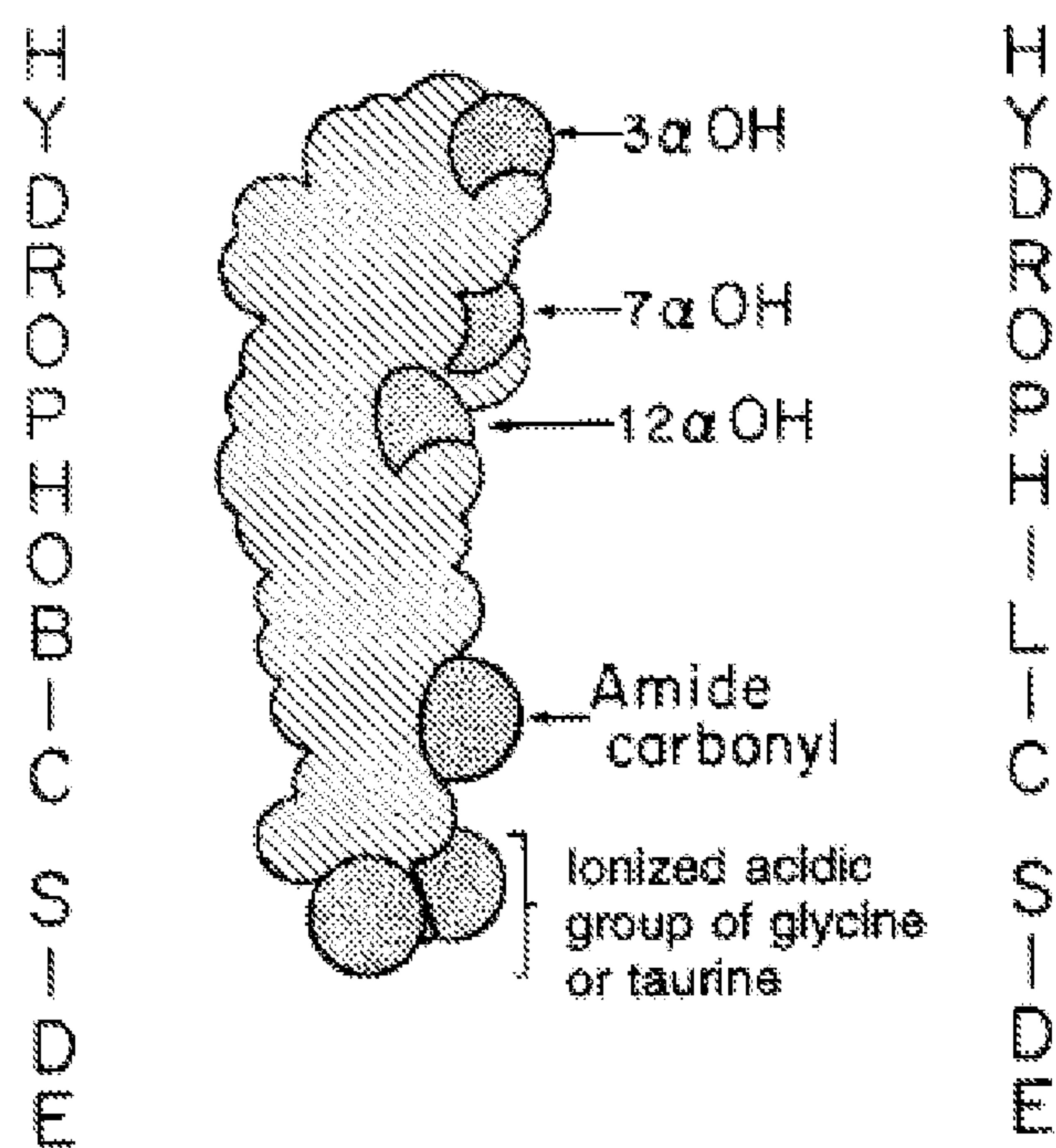
wherein said DCA has a  $^{14}\text{C}$  content of 1 ppt.



11. The compound of any one claims 8-10, which is the salt of the DCA.
12. The compound of claim 11 which is a sodium salt of the DCA.
13. A composition comprising the compound of any one of claims 8-10 and a pharmaceutically acceptable excipient.
14. The composition of claim 13 wherein the compound is a sodium salt of the DCA.
15. A method for determining whether a sample of deoxycholic acid was prepared synthetically, which method comprises:
  - a) assaying a sample of deoxycholic acid or a salt thereof for its fossil carbon or  $^{14}\text{C}$  content;
  - b) comparing the fossil carbon or  $^{14}\text{C}$  content of the sample against the fossil carbon or  $^{14}\text{C}$  content of naturally occurring deoxycholic acid to determine whether the sample was prepared synthetically.

1/2

FIGURE 1



**2/2**

FIGURE 2

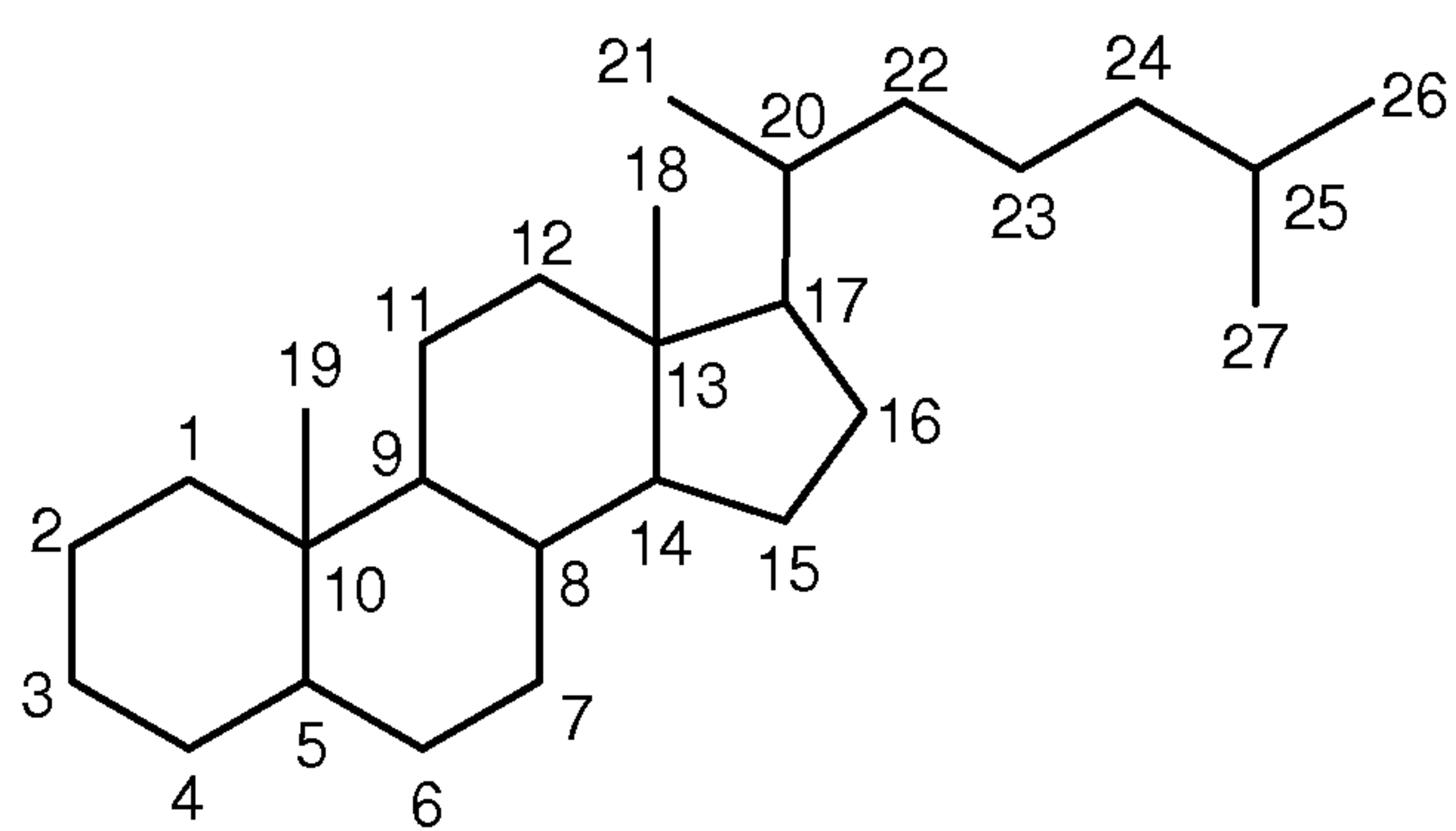
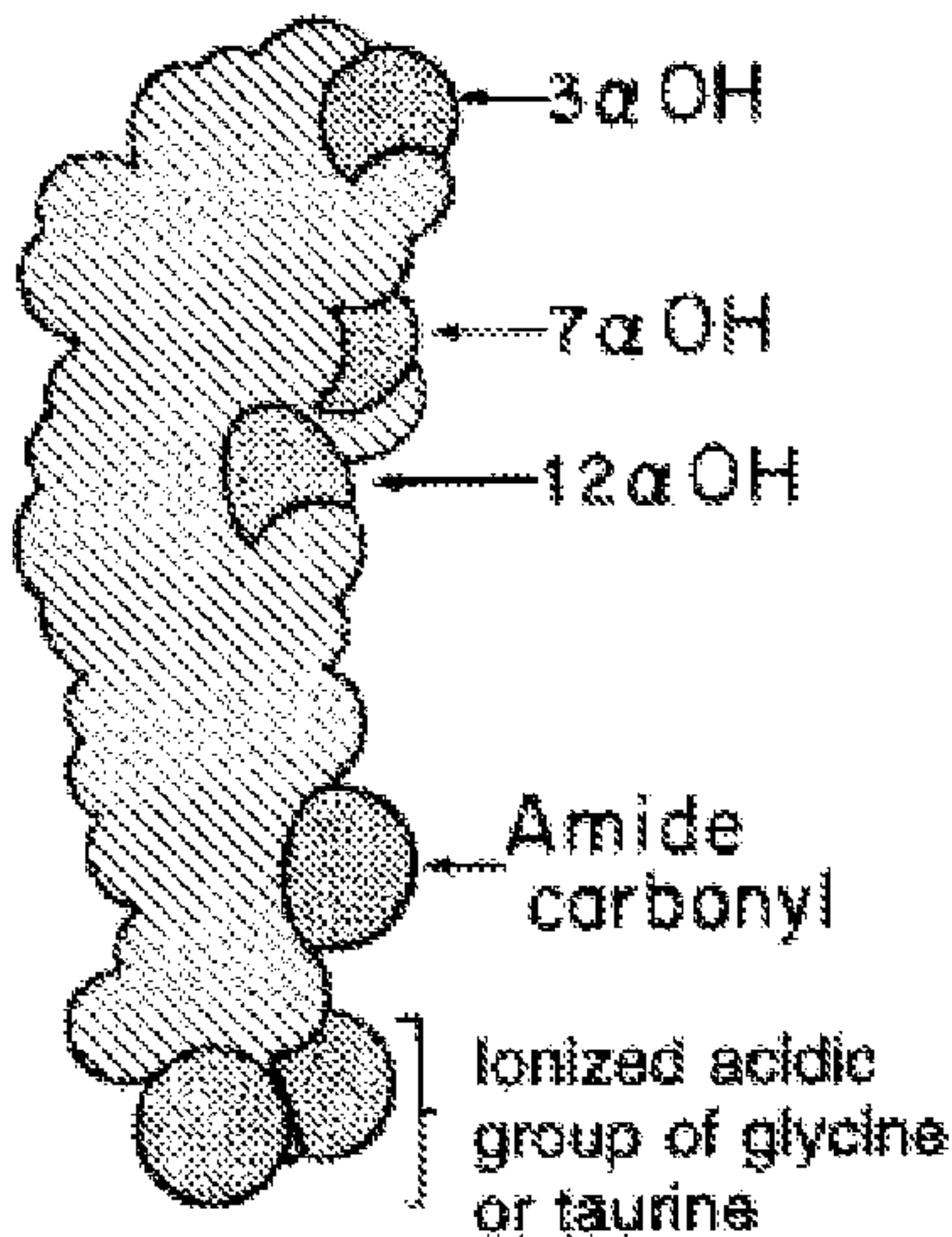




FIGURE 1

HYDROPHOBIC-C-S-MD



HYDROPHILIC-C-S-MD