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(54) Title: METALLOPHOSPHATE MOLECULAR SIEVES, METHOD OF PREPARATION AND USE

(57) Abstract: A new family of crystalline microporous metallophosphates designated AIPO-67 has been synthesized. These metallophosphates are represented by the empirical formula  $R^+M_m^{2+}E_pSi_nO_z$  where R is an organoammonium cation such as the ETMA<sup>+</sup> or DEDMA<sup>+</sup>, M is a framework metal alkaline earth or transition metal of valence 2+, and E is a trivalent framework element such as aluminum or gallium. The AIPO-67 compositions exhibit the LEV framework topology and have catalytic properties for carrying out various hydrocarbon conversion processes and separation properties for separating at least one component. The materials may be used as a catalytic composite for hydrocarbon conversions or may be used to separate at least one component from a mixture.



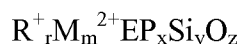
METALLOPHOSPHATE MOLECULAR SIEVES,  
METHOD OF PREPARATION AND USE

STATEMENT OF PRIORITY

[0001] This claims priority to U.S. Application No. 13/690,087 which was filed on  
5 November 30, 2012, and U.S. Application No. 13/690,078 which was filed on November 30,  
2012, now US 8,569,558, the contents of which are hereby incorporated by reference in its  
entirety.

FIELD OF THE INVENTION

[0002] This invention relates to a new family of charged metallophosphate-based  
10 molecular sieves designated AIPO-67, which has the LEV topology which may be used as a  
catalytic composite for hydrocarbon conversion reactions or as a separation material for  
separation processes. They are represented by the empirical formula of:



15 where M is a divalent framework metal such as magnesium or zinc, R is an  
organoammonium cation such as ethyltrimethylammonium and dimethylammonium and E is  
a trivalent framework element such as aluminum or gallium.

BACKGROUND OF THE INVENTION

[0003] Zeolites are crystalline aluminosilicate compositions which are microporous and  
20 which are formed from corner sharing  $AlO_2^-$  and  $SiO_2$  tetrahedra. Numerous zeolites, both  
naturally occurring and synthetically prepared are used in various industrial processes.  
Synthetic zeolites are prepared via hydrothermal synthesis employing suitable sources of Si,  
Al and structure directing agents such as alkali metals, alkaline earth metals, amines, or  
organoammonium cations. The structure directing agents reside in the pores of the zeolite  
25 and are largely responsible for the particular structure that is ultimately formed. These  
species balance the framework charge associated with aluminum and can also serve as space

fillers. Zeolites are characterized by having pore openings of uniform dimensions, having a significant ion exchange capacity, and being capable of reversibly desorbing an adsorbed phase which is dispersed throughout the internal voids of the crystal without significantly displacing any atoms which make up the permanent zeolite crystal structure. Zeolites can be used as catalysts for hydrocarbon conversion reactions, which can take place on outside surfaces of the zeolite as well as on internal surfaces within the pores of the zeolite.

[0004] In 1982, Wilson et al. developed aluminophosphate molecular sieves, typically called AIPOs, which are microporous materials that have many of the same properties of zeolites, but are silica free, composed of  $\text{AlO}_2^-$  and  $\text{PO}_2^+$  tetrahedra, see US 4310440.

Subsequently, charge was introduced to the neutral aluminophosphate frameworks via the substitution of  $\text{SiO}_2$  tetrahedra for  $\text{PO}_2^+$  tetrahedra to produce the SAPO molecular sieves, see US 4440871. Another way to introduce framework charge to neutral aluminophosphates is to substitute  $[\text{M}^{2+}\text{O}_2]^{2-}$  tetrahedra for  $\text{AlO}_2^-$  tetrahedra, which yield the MeAPO molecular sieves, see US 4567029. It is furthermore possible to introduce framework charge on AIPO-based molecular sieves via the introduction both of  $\text{SiO}_2$  and  $[\text{M}^{2+}\text{O}_2]^{2-}$  tetrahedra to the framework, giving MeAPSO molecular sieves, as shown in US 4973785.

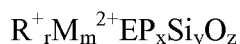
[0005] In these ground-breaking patents, metallophosphate materials with the LEV topology are disclosed. MAPO-35, a magnesium aluminophosphate material with the LEV topology is disclosed in US 4567029 in which quinuclidine is employed as a structure directing agent. Likewise, US 4440871 discloses SAPO-35, a silicoaluminophosphate that is also synthesized using the quinuclidine structure directing agent. In US 4973785, the MeAPSO composition CoAPSO-35 is disclosed, which contains both cobalt and silicon in the framework in addition to Al and P and uses methylquinuclidine as the structure directing agent. After this early work, SAPO-35 was prepared using cyclohexylamine as a structure directing agent, see Lohse et. al, *Crystal Research and Technology* (1993), 28(8), 1101-1107. Subsequently, SAPO-35 has been prepared using hexamethyleneimine as a structure directing agent under a variety of conditions, including in ethylene glycol, see Venkatathri et al., *JCS Faraday Transactions* (1997), 93(18), 3411-3415) and in aqueous and fluoride media, see Prakash et al, *Chem. Mater.* (1998), 10, 932-941. Wang et al. have recently disclosed MAPO-35 materials containing Mg and synthesized using the N-methylpiperidine and 1, 2-diaminocyclohexane as structure directing agents, see *Dalton Transactions* (2012), 41(22), 6855-6860). Most recently Cao et al. have disclosed the

preparation of SAPO-35 with the triethylmethylammonium structure directing agent, see US Pat. Appl. Publ. (2012), US 20120157741 A1.

[0006] In contrast to the above prior art, applicants have synthesized a variety of new materials and new structures using the very simple and commercially available ethyltrimethylammonium (ETMA<sup>+</sup>) and diethyldimethylammonium (DEDMA<sup>+</sup>) structure directing agents. One outcome of this investigation presented here is a new family of charged metallophosphate framework materials that contain a +3 valence metal, such as aluminum or gallium, and additionally at least one of a +2 valence metal, such as, for example, magnesium or zinc, and silicon, designated AIPO-67. When the +3 valence metal is Al, this corresponds to SAPO, MeAPO, and MeAPSO compositions. The AIPO-67 materials have the LEV topology that falls in the class of structures known as ABC-6 nets, see American Minerologist, 66, 777-788 (1981). The microporous AIPO-67 materials can be prepared with the simple ethyltrimethylammonium (ETMA<sup>+</sup>) and diethyldimethylammonium (DEDMA<sup>+</sup>) structure directing agents.

#### SUMMARY OF THE INVENTION

[0007] As stated, the present invention relates to a new family of metallophosphate molecular sieves designated AIPO-67. One embodiment of the invention is a microporous crystalline material having a three-dimensional framework of at least EO<sub>2</sub><sup>-</sup> and PO<sub>2</sub><sup>+</sup> tetrahedral units and furthermore, at least one of [M<sup>2+</sup>O<sub>2</sub>]<sup>2-</sup> and SiO<sub>2</sub> tetrahedral units and an empirical composition in the as synthesized and anhydrous basis expressed by an empirical formula of:



where M is at least one metal cation of valence +2 selected from the group consisting of Be<sup>2+</sup>, Mg<sup>2+</sup>, Zn<sup>2+</sup>, Co<sup>2+</sup>, Mn<sup>2+</sup>, Fe<sup>2+</sup>, Ni<sup>2+</sup>, "m" is the mole ratio of M to E and varies from 0 to 1.0, R is an organoammonium cation selected from the group consisting of ethyltrimethylammonium (ETMA<sup>+</sup>), diethyldimethylammonium (DEDMA<sup>+</sup>) and mixtures thereof, "r" is the mole ratio of R to E and has a value of 0.1 to 2.0, E is an +3 valence element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, "x" is mole ratio of P to E and varies from 0.5 to 2.0, "y" is the mole ratio of Si to E

and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (2 \cdot m + r + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings

5 and intensities set forth in Table A1:

Table A1

2 $\theta$	d(Å)	I/I <sub>0</sub> %
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s
13.65 - 13.16	6.48 - 6.72	m - vs
17.72 - 16.91	5.00 - 5.24	m - vs
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.5	3.96 - 4.13	s - vs
23.84 - 22.84	3.73 - 3.89	w - vs
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m
36.65 - 35.31	2.45 - 2.54	w - m
52.23 - 50.67	1.75 - 1.80	w - m

and in another embodiment, the material may be characterized by an x-ray pattern having at least the d-spacings and intensities set forth in Table A2:

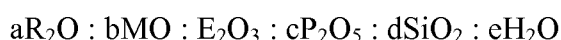
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Table A2

2 $\theta$	d(Å)	I/I <sub>0</sub> %
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s
11.90 - 11.44	7.43 - 7.73	w
13.65 - 13.16	6.48 - 6.72	m - vs
16.31 - 15.64	5.43 - 5.66	w - m
17.72 - 16.91	5.00 - 5.24	m - vs
18.09 - 17.51	4.90 - 5.06	w - m
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.50	3.96 - 4.13	s - vs
23.84 - 22.84	3.73 - 3.89	w - vs

25.43 - 24.57	3.50 - 3.62	w
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
29.76 - 28.59	3.00 - 3.12	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m
36.65 - 35.31	2.45 - 2.54	w - m
42.40 - 41.38	2.13 - 2.18	w
49.21 - 47.83	1.85 - 1.90	w
52.23 - 50.67	1.75 - 1.80	w - m

**[0008]** Another embodiment of the invention is a process for preparing the crystalline microporous metallophosphate molecular sieve described above. The process comprises forming a reaction mixture containing reactive sources of R, E, P, one or both of M and Si, and heating the reaction mixture at a temperature of 60°C to 200°C for a time sufficient to form the molecular sieve, the reaction mixture having a composition expressed in terms of mole ratios of the oxides of:



where “a” has a value of 0.25 to 16, “b” has a value of 0 to 2, “c” has a value of 0.8 to 8, “d” has a value of 0 to 4, and “e” has a value from 30 to 800, and wherein the sum of the values of “b” and “d” is greater than 0.01.

**[0009]** Yet another embodiment of the invention is a hydrocarbon conversion process using the above-described molecular sieve as a catalyst. The process comprises contacting at least one hydrocarbon with the molecular sieve at conversion conditions to generate at least one converted hydrocarbon.

**[0010]** Still another embodiment of the invention is a separation process using the crystalline AlPO-67 material. The process may involve separating mixtures of molecular species or removing contaminants by contacting a fluid with the AlPO-67 molecular sieve. Separation of molecular species can be based either on the molecular size (kinetic diameter) or on the degree of polarity of the molecular species. Removing contaminants may be by ion exchange with the molecular sieve.

## DETAILED DESCRIPTION OF THE INVENTION

[0011] Applicants have prepared a family of metallophosphate materials with the ETMA<sup>+</sup> and DEDMA<sup>+</sup> structure directing agents whose topological structure is LEV, see [http://izasc-mirror.la.asu.edu/fmi/xsl/IZA-SC/ftc\\_fw.xsl?-db=Atlas\\_main&-lay=fw&-](http://izasc-mirror.la.asu.edu/fmi/xsl/IZA-SC/ftc_fw.xsl?-db=Atlas_main&-lay=fw&-max=25&STC=LEV&-find)

5 max=25&STC=LEV&-find. In their paper “Enumeration of 4-connected 3-dimensional nets and classification of framework silicates: the infinite set of ABC-6 nets; the Archimedean and  $\sigma$ -related nets,” Smith and Bennett state “To a first approximation, all silicates belonging to the ABC-6 net family have x-ray diffraction patterns which can be indexed on a hexagonal prismatic unit cell with lattice parameters  $a = 13.0 \pm 0.3 \text{ \AA}$  and  $c = p \times (2.6 \pm 0.1 \text{ \AA})$ ,” see  
 10 American Mineralogist, 66, 777-788 (1981). The AlPO-67 materials index on such hexagonal unit cells as one SAPO-67 material has lattice parameters  $a = 13.155 \text{ \AA}$  and  $c = 22.813 \text{ \AA}$ , which suggests an ABC-6 net structure with the stacking sequence repeating every 9 layers along the c-axis ( $p = 22.813/2.6 = 9.12$ ), a characteristic of the LEV topology. AlPO-67 is the fourth ABC 6-net topology isolated in our investigation of the structure directing properties of ETMA<sup>+</sup> and  
 15 DEDMA<sup>+</sup>. The instant microporous crystalline material (AlPO-67) has an empirical composition in the as-synthesized form and on an anhydrous basis expressed by the empirical formula:

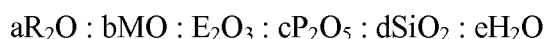


where M is at least one divalent cation and is selected from the group consisting of alkaline earth  
 20 and transition metals. Specific examples of the M cations include but are not limited to beryllium, magnesium, cobalt (II), manganese, zinc, iron(II), nickel and mixtures thereof. R is an organoammonium cation, examples of which include but are not limited to ethyltrimethylammonium (ETMA<sup>+</sup>), and diethyldimethylammonium (DEDMA<sup>+</sup>) and mixtures thereof and “r” is the mole ratio of R to E and varies from 0.1 to 2.0. The value of “m” is the  
 25 mole ratio of M to E and varies from 0 to 1.0, “x” is mole ratio of P to E and varies from 0.5 to 2.0. The ratio of silicon to E is represented by “y” which varies from 0 to 1.0 and “m” + “y”  $\geq$  0.02. E is an trivalent element which is tetrahedrally coordinated, is present in the framework and is selected from the group consisting of aluminum, gallium, iron(III) and boron. Lastly, “z” is the mole ratio of O to E and is given by the equation:

$$z = (2 \bullet m + r + 3 + 5 \bullet x + 4 \bullet y)/2.$$

[0012] The microporous crystalline metallophosphate AIPO-67 is prepared by a hydrothermal crystallization of a reaction mixture prepared by combining reactive sources of R, E, phosphorous, and one or both of M and silicon. A preferred form of the AIPO-67 materials is when E is Al. The sources of aluminum include but are not limited to aluminum alkoxides, precipitated aluminas, aluminum metal, aluminum hydroxide, aluminum salts and alumina sols. Specific examples of aluminum alkoxides include, but are not limited to aluminum ortho sec-butoxide and aluminum ortho isopropoxide. Sources of the other E elements include but are not limited to organoammonium borates, boric acid, precipitated gallium oxyhydroxide, gallium sulfate, ferric sulfate, and ferric chloride. Sources of phosphorus include, but are not limited to, orthophosphoric acid, phosphorus pentoxide, and ammonium dihydrogen phosphate. Sources of silica include but are not limited to tetraethylorthosilicate, colloidal silica, and precipitated silica. Sources of the M metals include the halide salts, nitrate salts, acetate salts, and sulfate salts of the respective alkaline earth and transition metals. R is an organoammonium cation selected from the group consisting of  $\text{ETMA}^+$  and  $\text{DEDMA}^+$ , and mixtures thereof, and the sources include the hydroxide, chloride, bromide, iodide and fluoride compounds. Specific examples include without limitation ethyltrimethylammonium hydroxide, ethyltrimethylammonium chloride, diethyltrimethylammonium chloride and diethyltrimethylammonium hydroxide. In one embodiment R is  $\text{ETMA}^+$ . In another embodiment R is  $\text{DEDMA}^+$ . In yet another embodiment, R is a combination of  $\text{ETMA}^+$  and  $\text{DEDMA}^+$ .

[0013] The reaction mixture containing reactive sources of the desired components can be described in terms of molar ratios of the oxides by the formula:



where "a" varies from 0.25 to 16, "b" varies from 0 to 2, "c" varies from 0.8 to 8, "d" varies from 0 to 4, and "e" varies from 30 to 800, and wherein the sum of the values of "b" and "d" is greater than 0.01. If alkoxides are used, a distillation or evaporative step may be included to remove the alcohol hydrolysis products. The reaction mixture is now reacted at a temperature of 60°C to 200°C, and in another embodiment from 125°C to 175°C, for a period of 1 day to 3 weeks, and in another embodiment for a period of 2 days to 10 days, in a sealed reaction vessel at autogenous pressure. After crystallization is complete, the solid product is isolated from the heterogeneous mixture by means such as filtration or centrifugation, and then washed with deionized water and dried in air at ambient temperature up to 100°C. AIPO-67 seeds can



optionally be added to the reaction mixture in order to accelerate the formation of the desired microporous composition.

[0014] The AlPO-67 metallophosphate-based material, which is obtained from the above-described process, is characterized by the x-ray diffraction pattern, having at least the d-spacings and relative intensities set forth in Table A1 below.

Table A1

$2\Theta$	$d(\text{\AA})$	$I/I_0\%$
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s
13.65 - 13.16	6.48 - 6.72	m - vs
17.72 - 16.91	5.00 - 5.24	m - vs
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.5	3.96 - 4.13	s - vs
23.84 - 22.84	3.73 - 3.89	w - vs
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m
36.65 - 35.31	2.45 - 2.54	w - m
52.23 - 50.67	1.75 - 1.80	w - m

[0015] In another embodiment of the invention, AlPO-67 metallophosphate, which is obtained from the above-described process, is characterized by the x-ray diffraction pattern, having at least the d-spacings and relative intensities set forth in Table A2 below.

Table A2

$2\Theta$	$d(\text{\AA})$	$I/I_0\%$
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s
11.90 - 11.44	7.43 - 7.73	w
13.65 - 13.16	6.48 - 6.72	m - vs
16.31 - 15.64	5.43 - 5.66	w - m
17.72 - 16.91	5.00 - 5.24	m - vs
18.09 - 17.51	4.90 - 5.06	w - m
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.50	3.96 - 4.13	s - vs

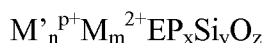
23.84 - 22.84	3.73 - 3.89	w - vs
25.43 - 24.57	3.50 - 3.62	w
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
29.76 - 28.59	3.00 - 3.12	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m
36.65 - 35.31	2.45 - 2.54	w - m
42.40 - 41.38	2.13 - 2.18	w
49.21 - 47.83	1.85 - 1.90	w
52.23 - 50.67	1.75 - 1.80	w - m

[0016] In one embodiment of the invention, the AlPO-67 is thermally stable up to a temperature of at least 400°C, and in another embodiment the AlPO-67 is thermally stable up to a temperature of at least 500°C.

5 [0017] The AlPO-67 may be modified in many ways to tailor it for use in a particular application. Modifications include calcination, ammonia calcinations, ion-exchange, steaming, various acid extractions, ammonium hexafluorosilicate treatment, or any combination thereof, as outlined for the case of UZM-4 in US 6,776,975 B1 which is incorporated by reference in its entirety. Properties that may be modified include porosity,  
10 adsorption, framework composition, acidity, thermal stability, etc.

[0018] As synthesized, the AlPO-67 material will contain some of the exchangeable or charge balancing cations in its pores. These exchangeable cations can be exchanged for other cations, or in the case of organic cations, they can be removed by heating under controlled conditions. In one embodiment, the method of removing organic cations from the pores is  
15 ammonia calcination. Calcination in air converts the organic cations in the pores to protons, which can, for example, lead to some removal of Al from the framework upon exposure to water vapor. When the calcination is carried out in an ammonia atmosphere, the organic cation in the pore is replaced by NH<sub>4</sub><sup>+</sup> cation and the framework remains intact, see Studies in Surface Science, (2004) vol. 154, p.1324 - 1331. Typical conditions for ammonia calcinations  
20 include the use of gaseous anhydrous ammonia flowing at a rate of 1.1 l/min while ramping the sample at 2 -5°C/min to 500°C and holding at that temperature for a time ranging from 5 minutes to 5 hours. The resulting ammonium form of AlPO-67 has essentially the diffraction pattern of Table A1. The ammonium form of AlPO-67 may then be ion-exchanged to any

other form, resulting in a material with a modified composition, AlPO-67M, given by the empirical formula:



where M is at least one framework metal cation of valence +2 selected from the group consisting of Be<sup>2+</sup>, Mg<sup>2+</sup>, Zn<sup>2+</sup>, Co<sup>2+</sup>, Mn<sup>2+</sup>, Fe<sup>2+</sup>, Ni<sup>2+</sup>, “m” is the mole ratio of M to E and varies from 0 to 1.0, M’ is selected from the group consisting of NH<sub>4</sub><sup>+</sup>, H<sup>+</sup>, alkali metals, alkaline earth metals, rare earth metals and mixtures thereof, “n” is the mole ratio of M’ to E and has a value of 0.03 to 2.0, “p” is the weighted average valence of M’ and varies from 1 to 3, E is a trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y” ≥ 0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (p \cdot n + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

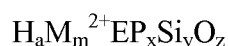
and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1:

Table A1

2 $\Theta$	d(Å)	I/I <sub>0</sub> %
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s
13.65 - 13.16	6.48 - 6.72	m - vs
17.72 - 16.91	5.00 - 5.24	m - vs
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.5	3.96 - 4.13	s - vs
23.84 - 22.84	3.73 - 3.89	w - vs
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m
36.65 - 35.31	2.45 - 2.54	w - m
52.23 - 50.67	1.75 - 1.80	w - m

[0019] When AlPO-67 is calcined in air, there can be loss of metal from the framework, such as Al, which can alter the x-ray diffraction pattern from that observed for the as-synthesized AlPO-67, see Studies in Surface Science, (2004) vol. 154, p.1324 - 1331. Some

AlPO-67 compositions may not be stable to air calcination and subsequent exposure to water. The stable air calcined AlPO-67 materials, AlPO-67C, are characterized on an anhydrous basis by the empirical formula:



5 where M is at least one metal cation of valence +2 selected from the group consisting of Be<sup>2+</sup>, Mg<sup>2+</sup>, Zn<sup>2+</sup>, Co<sup>2+</sup>, Mn<sup>2+</sup>, Fe<sup>2+</sup>, Ni<sup>2+</sup>, “m” is the mole ratio of M to E and varies from 0 to 1.0, H is a proton, “a” is the mole ratio of H to E and has a value of 0.1 to 2.0, E is an trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E  
10 and varies from 0.0 to 1.0, “m” + “y” ≥ 0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (a + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table B:

15

Table B

2 $\Theta$	d(Å)	I/I <sub>0</sub> %
9.02 - 8.33	9.80 - 10.6	w - m
11.09 - 10.79	7.97 - 8.19	vs
13.74 - 13.32	6.44 - 6.64	s - vs
17.58 - 16.97	5.04 - 5.22	m
21.19 - 20.69	4.19 - 4.29	m
22.32 - 21.87	3.98 - 4.06	m - s
28.78 - 28.13	3.10 - 3.17	m
32.78 - 32.05	2.73 - 2.79	w - m

[0020] The crystalline AlPO-67 materials of this invention can be used for separating mixtures of molecular species, removing contaminants through ion exchange and catalyzing various hydrocarbon conversion processes. Separation of molecular species can be based either  
20 on the molecular size (kinetic diameter) or on the degree of polarity of the molecular species. The separation involves contacting the mixture to be separated with the AlPO-67 material. Further processing may be involved.

[0021] The AlPO-67 compositions of this invention can also be used as a catalyst or catalyst support in various hydrocarbon conversion processes. Hydrocarbon conversion processes are well known in the art and include cracking, hydrocracking, alkylation of both aromatics and isoparaffin, isomerization, polymerization, reforming, hydrogenation, dehydrogenation, 5 transalkylation, dealkylation, hydration, dehydration, hydrotreating, hydrodenitrogenation, hydrodesulfurization, methanol to olefins, methanation and syngas shift process. Specific reaction conditions and the types of feeds which can be used in these processes are set forth in US 4,310,440, US 4,440,871 and US 5,126,308, which are hereby incorporated by reference. Preferred hydrocarbon conversion processes are those in which hydrogen is a component such 10 as hydrotreating or hydrofining, hydrogenation, hydrocracking, hydrodenitrogenation, hydrodesulfurization, etc.

[0022] Hydrocracking conditions typically include a temperature in the range of 400° to 1200°F (204-649°C), preferably between 600° and 950°F (316-510°C). Reaction pressures are in the range of atmospheric to 3,500 psig (24,132 kPa g), preferably between 200 and 3000 psig 15 (1379 – 20,685 kPa g). Contact times usually correspond to liquid hourly space velocities (LHSV) in the range of 0.1 hr<sup>-1</sup> to 15 hr<sup>-1</sup>, preferably between 0.2 and 3 hr<sup>-1</sup>. Hydrogen circulation rates are in the range of 1,000 to 50,000 standard cubic feet (scf) per barrel of charge (178-8,888 std. m<sup>3</sup>/m<sup>3</sup>), preferably between 2,000 and 30,000 scf per barrel of charge (355-5,333 std. m<sup>3</sup>/m<sup>3</sup>). Suitable hydrotreating conditions are generally within the broad ranges of 20 hydrocracking conditions set out above.

[0023] The reaction zone effluent is normally removed from the catalyst bed, subjected to partial condensation and vapor-liquid separation and then fractionated to recover the various components thereof. The hydrogen, and if desired some or all of the unconverted heavier materials, are recycled to the reactor. Alternatively, a two-stage flow may be employed with the 25 unconverted material being passed into a second reactor. Catalysts of the subject invention may be used in just one stage of such a process or may be used in both reactor stages.

[0024] Catalytic cracking processes are preferably carried out with the AlPO-67 composition using feedstocks such as gas oils, heavy naphthas, deasphalted crude oil residua, etc. with gasoline being the principal desired product. Temperature conditions of 850° to 30 1100°F, LHSV values of 0.5 to 10 and pressure conditions of from 0 to 50 psig are suitable.

[0025] Alkylation of aromatics usually involves reacting an aromatic (C<sub>2</sub> to C<sub>12</sub>), especially benzene, with a monoolefin to produce a linear alkyl substituted aromatic. The process is

carried out at an aromatic: olefin (e.g., benzene: olefin) ratio of between 5:1 and 30:1, a LHSV of 0.3 to 6 hr<sup>-1</sup>, a temperature of 100° to 250°C and pressures of 200 to 1000 psig. Further details on apparatus may be found in US 4,870,222 which is incorporated by reference.

**[0026]** Alkylation of isoparaffins with olefins to produce alkylates suitable as motor fuel components is carried out at temperatures of -30° to 40°C, pressures from atmospheric to 6,894 kPa (1,000 psig) and a weight hourly space velocity (WHSV) of 0.1 to 120. Details on paraffin alkylation may be found in US 5,157,196 and US 5,157,197, which are incorporated by reference

**[0027]** The conversion of methanol to olefins is effected by contacting the methanol with the AlPO-67 catalyst at conversion conditions, thereby forming the desired olefins. The methanol can be in the liquid or vapor phase with the vapor phase being preferred. Contacting the methanol with the AlPO-67 catalyst can be done in a continuous mode or a batch mode with a continuous mode being preferred. The amount of time that the methanol is in contact with the AlPO-67 catalyst must be sufficient to convert the methanol to the desired light olefin products. When the process is carried out in a batch process, the contact time varies from 0.001 hrs to 1 hr and preferably from 0.01 hr to 1.0 hr. The longer contact times are used at lower temperatures while shorter times are used at higher temperatures. Further, when the process is carried out in a continuous mode, the Weight Hourly Space Velocity (WHSV) based on methanol can vary from 1 hr<sup>-1</sup> to 1000 hr<sup>-1</sup> and preferably from 1hr<sup>-1</sup> to 100 hr<sup>-1</sup>.

**[0028]** Generally, the process must be carried out at elevated temperatures in order to form light olefins at a fast enough rate. Thus, the process should be carried out at a temperature of 300 °C to 600 °C, preferably from 400 °C to 550 °C and most preferably from 450 °C to 525 °C. The process may be carried out over a wide range of pressure including autogenous pressure. Thus, the pressure can vary from 0 kPa (0 psig) to 1724 kPa (250 psig) and preferably from 34 kPa (5 psig) to 345 kPa (50 psig).

**[0029]** Optionally, the methanol feedstock may be diluted with an inert diluent in order to more efficiently convert the methanol to olefins. Examples of the diluents which may be used are helium, argon, nitrogen, carbon monoxide, carbon dioxide, hydrogen, steam, paraffinic hydrocarbons, e. g., methane, aromatic hydrocarbons, e. g., benzene, toluene and mixtures thereof. The amount of diluent used can vary considerably and is usually from 5 to 90 mole percent of the feedstock and preferably from 25 to 75 mole percent.

[0030] The actual configuration of the reaction zone may be any known catalyst reaction apparatus known in the art. Thus, a single reaction zone or a number of zones arranged in series or parallel may be used. In such reaction zones the methanol feedstock is flowed through a bed containing the AlPO-67 catalyst. When multiple reaction zones are used, one or more AlPO-67 catalysts may be used in series to produce the desired product mixture. Instead of a fixed bed, a dynamic bed system, e. g., fluidized or moving, may be used. Such a dynamic system would facilitate any regeneration of the AlPO-67 catalyst that may be required. If regeneration is required, the AlPO-67 catalyst can be continuously introduced as a moving bed to a regeneration zone where it can be regenerated by means such as oxidation in an oxygen containing atmosphere to remove carbonaceous materials.

[0031] The following examples are presented in illustration of this invention and are not intended as undue limitations on the generally broad scope of the invention as set out in the appended claims. The products will be designated with names that contain the suffix “-67” to indicate the “-67” structure and prefix that reflects the compositional nature of the product, such as “SAPO” for a silicoaluminophosphate, MAPO for a metalloaluminophosphate, and MAPSO for a silicometalloaluminophosphate, etc.

[0032] The structure of the AlPO-67 compositions of this invention was determined by x-ray analysis. The x-ray patterns presented in the following examples were obtained using standard x-ray powder diffraction techniques. The radiation source was a high-intensity, x-ray tube operated at 45 kV and 35 mA. The diffraction pattern from the copper K-alpha radiation was obtained by appropriate computer based techniques. Flat compressed powder samples were continuously scanned at  $2^{\circ}$  to  $56^{\circ}$  ( $2\theta$ ). Interplanar spacings ( $d$ ) in Angstrom units were obtained from the position of the diffraction peaks expressed as  $\theta$  where  $\theta$  is the Bragg angle as observed from digitized data. Intensities were determined from the integrated area of diffraction peaks after subtracting background, “ $I_0$ ” being the intensity of the strongest line or peak, and “ $I$ ” being the intensity of each of the other peaks.

[0033] As will be understood by those skilled in the art the determination of the parameter  $2\theta$  is subject to both human and mechanical error, which in combination can impose an uncertainty of  $\pm 0.4^{\circ}$  on each reported value of  $2\theta$ . This uncertainty is, of course, also manifested in the reported values of the  $d$ -spacings, which are calculated from the  $2\theta$  values. This imprecision is general throughout the art and is not sufficient to preclude the differentiation of

the present crystalline materials from each other and from the compositions of the prior art. In some of the x-ray patterns reported, the relative intensities of the d-spacings are indicated by the notations vs, s, m, and w which represent very strong, strong, medium, and weak, respectively. In terms of  $100 \times I/I_0$ , the above designations are defined as:

5  $w = 0-15; m = 15-60; s = 60-80$  and  $vs = 80-100$

[0034] In certain instances the purity of a synthesized product may be assessed with reference to its x-ray powder diffraction pattern. Thus, for example, if a sample is stated to be pure, it is intended only that the x-ray pattern of the sample is free of lines attributable to crystalline impurities, not that there are no amorphous materials present.

10 [0035] In order to more fully illustrate the invention, the following examples are set forth. It is to be understood that the examples are only by way of illustration and are not intended as an undue limitation on the broad scope of the invention as set forth in the appended claims.

#### EXAMPLE 1

15 [0036] A Teflon beaker was charged with 15.15 g  $H_3PO_4$  (85%) to which 13.23 g deionized water was added with stirring. Separately, 5.00 g  $Zn(OAc)_2 \cdot 2H_2O$  was dissolved in 50.00 g de-ionized water. This solution was added to the  $H_3PO_4$  with vigorous stirring. Next, 11.1 g  $Al(OH)_3$  (80.0%) was added slowly and intermittently to allow the formation of a milky suspension. After the addition was completed, the reaction mixture was homogenized  
 20 for 2 hours. Over a period of 3 minutes, 40.75 g DEDMAOH (20 %) was added in 5 pours. The mixture was stirred for 20 minutes. Portions of the reaction mixture were distributed among 7 Teflon-lined autoclaves and digested at autogenous pressure at range of temperatures for different time periods. The products were isolated by centrifugation and washed with de-ionized water, and dried at room temperature. The products from the 100 °C  
 25 and 125 °C reactions that were digested for 36 hours were identified as AlPO-67 via powder x-ray diffraction. The diffraction lines representative of the 125 °C product are shown below in Table 1.

Table 1

2- $\Theta$	d(Å)	I/I <sub>0</sub> %
8.70	10.16	m
11.00	8.04	s



11.78	7.50	w
13.38	6.61	m
15.98	5.54	w
17.38	5.10	s
17.88	4.96	w
21.10	4.21	m
21.94	4.05	vs
23.28	3.82	m
25.16	3.54	w
26.90	3.31	m
28.56	3.12	m
29.39	3.04	w
32.12	2.78	m
34.48	2.60	w
35.82	2.50	w
42.02	2.15	w
42.40	2.13	w
42.78	2.11	w
44.62	2.03	w
48.54	1.87	w
51.32	1.78	w

## EXAMPLE 2

[0037] A Teflon beaker was charged with 97.76 g DEDMAOH (20 %) and placed under a high speed mixer. To this solution, 4.60 g Al(OH)<sub>3</sub> (26.7 %) was added in three portions and this dissolved to form a clear solution. Next, 13.12 g H<sub>3</sub>PO<sub>4</sub> (85%) was added dropwise, resulting in a clear solution. Separately, 2.23 g Mn(OAc)<sub>2</sub>\*4H<sub>2</sub>O was dissolved in 10.00 g de-ionized water. This solution was added dropwise and immediately introduced some haziness to reaction mixture, so the addition was terminated. Then 1.01 g H<sub>3</sub>PO<sub>4</sub> (85%) was diluted with 1.00 g de-ionized water and added to the reaction mixture in a single pour. The reaction mixture was stirred until it was clear. Then the addition of the Mn acetate solution was continued. The reaction mixture was homogenized for an additional 30 minutes after completion of the addition, at which time it was a clear solution. The reaction mixture was distributed among several Teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and time periods. The products were isolated by centrifugation, washed with de-ionized water and dried at room temperature. The portion of the reaction

mixture digested at 175°C for 36 hours yielded MAPO-67, which was identified by powder x-ray diffraction. Representative lines in the diffraction pattern for the MAPO-67 product are given in Table 2.

Table 2

2- $\Theta$	d(Å)	I/I <sub>0</sub> %
8.58	10.30	w
10.86	8.14	m
11.58	7.64	w
13.28	6.66	m
15.81	5.60	w
17.20	5.15	S
17.66	5.02	W
20.82	4.26	M
21.80	4.07	V <sub>s</sub>
23.12	3.84	M
24.80	3.59	W
26.76	3.33	m
28.24	3.16	M
28.96	3.08	W
31.40	2.85	W
32.02	2.79	M
34.24	2.62	W
34.81	2.58	W
35.22	2.55	W
35.68	2.51	W
36.80	2.44	W
41.74	2.16	W
44.52	2.03	W
47.56	1.91	W
48.21	1.89	W
49.40	1.84	W
50.96	1.79	W
54.87	1.67	W
55.16	1.66	W

## EXAMPLE 3

[0038] A Teflon beaker was charged with 100.00 g DEDMAOH (20 %) which was placed under a high speed stirrer. Then 4.65 g Al(OH)<sub>3</sub> (78.1%) was added intermittently with stirring until it was all nearly dissolved. Phosphoric acid (85%), 19.34 g, was

5 continuously added dropwise until the addition was completed, during which time the reaction mixture became a clear solution. Separately, 2.05 g Zn(OAc)<sub>2</sub>\*2H<sub>2</sub>O was dissolved in 10.00 g de-ionized water. This solution was added dropwise to the reaction mixture, which was hazy but cleared up with homogenization. The reaction mixture was distributed among 7

10 Teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and time periods. The products were isolated by centrifugation, washed with de-ionized water, and dried at room temperature. Over the entire range of conditions examined, 95 °C to 175 °C for 61 – 159 hours, the reactions yielded MAPO-67 as the product, which were identified by powder x-ray diffraction. The characteristic diffraction lines for the 150°C/61 hour

15 product are given below in Table 3. Elemental analysis on this same product shows the composition to consist of the elemental ratios N/Al = 0.25, Zn/Al = 0.26, and P/Al = 1.22, while C/N = 6.32. A portion of this material was calcined at 500°C, using a ramp rate of 2°C/min in a nitrogen atmosphere, calcining in nitrogen for 4 hours before switching to dry air and calcining for an additional 2 hours at 500°C. The material was identified as AlPO-67C by powder x-ray diffraction. The diffraction lines for the calcined material are shown in table

20 3 below.

Table 3

Example 3			Example 3 calcined		
2- $\Theta$	d(Å)	I/I <sub>0</sub> %	2 $\Theta$	d(Å)	I/I <sub>0</sub> %
8.50	10.39	w	8.49	10.41	w
10.68	8.28	m	10.88	8.13	vs
13.28	6.66	m	13.62	6.50	vs
17.10	5.18	m	17.14	5.17	m
20.85	4.26	m	20.96	4.23	m

20.90	4.25	m	21.5	4.13	m
21.72	4.09	vs	22.14	4.01	m
23.10	3.85	m	28.34	3.15	m
26.82	3.32	m	32.54	2.75	m
28.70	3.11	m			
31.98	2.80	m			
32.69	2.74	w			
34.56	2.59	w			
35.60	2.52	w			
42.01	2.15	w			
42.21	2.14	w			
48.68	1.87	w			
51.16	1.78	w			
55.13	1.66	w			

## EXAMPLE 4

[0039] A Teflon beaker was charged with 130.00 g ETMAOH (20%) and placed under a high speed stirrer. Then 3.71 g LUDOX AS-40<sup>TM</sup> colloidal silica (40 % SiO<sub>2</sub>) was added, stirred briefly and the reaction mixture was transferred to Teflon bottle, sealed, and placed in an oven at 100°C for one hour. The resulting solution was returned to the Teflon beaker and 12.34 g Al(OH)<sub>3</sub> (78.1%) was added with vigorous stirring, stirring for an hour after the addition was completed. Then 35.62 g H<sub>3</sub>PO<sub>4</sub> (85%) was added dropwise and intermittently. At the end of the addition, the solution was nearly clear. After homogenizing for another hour, the reaction mixture was distributed among 8 teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and time periods. The products were isolated by centrifugation, washed with de-ionized water, and dried at room temperature. All of the products, digested under conditions ranging from 24 – 209 hours and 150 – 200°C, were identified as SAPO-67 via powder x-ray diffraction. Representative diffraction lines for the product from the 175 °C/109 hour digestion are given below in Table 4. Elemental analysis on this product showed the composition to consist of the elemental ratios N/Al = 0.17, Na/Al = 0.02 (from Ludox AS-40<sup>TM</sup>), Si/Al = 0.22, P/Al = 0.81, while C/N = 5.21.

Table 4

2- $\Theta$	d(Å)	I/I <sub>0</sub> %
8.64	10.22	w
10.96	8.07	m
11.63	7.60	w
13.40	6.60	m
15.98	5.54	w
17.36	5.11	s
17.79	4.98	w
20.96	4.24	m
21.98	4.04	vs
23.32	3.81	m
24.96	3.56	w
27.00	3.30	m
28.44	3.14	m
29.22	3.05	w
31.60	2.83	w
32.24	2.77	m
34.42	2.60	w
35.11	2.55	w
35.96	2.50	w
37.79	2.38	w
41.03	2.20	w
41.98	2.15	w
42.76	2.11	w
44.79	2.02	w
47.82	1.90	w
48.58	1.87	w
49.00	1.86	w
49.74	1.83	w
51.32	1.78	w
55.20	1.66	w
55.60	1.65	w

## EXAMPLE 5

**[0040]** A Teflon beaker was charged with 100.00 g ETMAOH (20 %) and placed under a high speed mixer. Then 7.61 g Ludox AS-40<sup>TM</sup> colloidal silica (40 % SiO<sub>2</sub>) was added to the reaction mixture, which was stirred briefly, transferred to a Teflon bottle, sealed, and digested in an oven at 95°C for 1 hour. The resulting solution was transferred back to the beaker,

placed under the stirrer, and 25.31 g Al(OH)<sub>3</sub> (78.1%) was added in small portions. After initially dissolving, a white suspension eventually formed. The reaction mixture was homogenized for 20 minutes after the addition. Separately, 43.84 g H<sub>3</sub>PO<sub>4</sub> (85%) was diluted with 22.83 g de-ionized water. This was added dropwise to the reaction mixture, 5 intermittently, allowing the mixture to stir between additions. At a certain point during the addition the reaction mixture got very thick and some hand mixing was required before going to high speed mixing while supporting the beaker. The remainder of the phosphoric acid solution was added while mixing vigorously. The resulting reaction mixture was a moist homogenous viscous white gel. The reaction mixture was distributed among 8 Teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and times. The products were isolated by centrifugation, washed with de-ionized water and dried at room temperature. X-ray diffraction analysis identified the products from 150 and 175°C digestions as SAPO-67, while the 200°C digestions started to show SAPO-67 with impurities after 2 days. Characteristic diffraction lines for the product resulting from 150°C/108 hour digestion are shown below in Table 5. A portion of this material was calcined at 500°C, using a ramp rate of 2°C/min in a nitrogen atmosphere, calcining in nitrogen for 4 hours before switching to dry air and calcining for an additional 2 hours at 500°C. The material was identified as AlPO-67C by powder x-ray diffraction. The diffraction lines for the calcined material are shown in table 5 below.

20

Table 5

Example 5			Example 5 calcined		
2- $\Theta$	d(Å)	I/I <sub>0</sub> %	2 $\Theta$	d(Å)	I/I <sub>0</sub> %
8.74	10.11	m	8.64	10.23	m
11.02	8.02	m	10.92	8.10	vs
11.70	7.56	w	11.57	7.64	w
13.52	6.54	m	13.42	6.59	vs
16.12	5.50	w	16.02	5.53	w
17.50	5.07	s	17.34	5.11	m
17.92	4.95	w	17.74	5.00	w

21.10	4.21	m	20.86	4.26	m
22.16	4.01	vs	22.00	4.04	m
23.54	3.78	m	23.38	3.80	w
25.12	3.54	w	24.90	3.57	w
27.22	3.27	m	27.08	3.29	w
28.66	3.11	m	28.36	3.14	m
29.45	3.03	w	29.24	3.05	w
31.82	2.81	w	31.46	2.84	w
32.54	2.75	m	32.32	2.77	m
34.70	2.58	w	34.36	2.61	w
35.36	2.54	w	35.18	2.55	w
36.26	2.48	w	36.12	2.49	w
42.36	2.13	w	42.02	2.15	w
43.10	2.10	w	42.72	2.12	w
44.64	2.03	w	44.90	2.02	w
45.20	2.00	w	47.54	1.91	w
48.16	1.89	w	48.58	1.87	w
49.05	1.86	w	49.24	1.85	w
50.23	1.81	w	49.96	1.82	w
51.80	1.76	w	51.30	1.78	w
55.68	1.65	w			

## EXAMPLE 6

[0041] A Teflon beaker was charged with 130.00 g DEDMAOH (20%) to which 7.46 g Ludox AS-40<sup>TM</sup> colloidal silica (40% silica) was added. The mixture was placed in a Teflon bottle, sealed and digested in an oven at 95°C for 1.5 hours. The resulting solution transferred to a beaker, placed under a high speed mixer, and with vigorous stirring, 44.63 g Al(OsecBu)<sub>3</sub> (97%) was added in 5 pours. The clear solution was stirred for an hour. Next,

25.15 g H<sub>3</sub>PO<sub>4</sub> (85%) was added dropwise and intermittently. The solution became warm and a little cloudy. After 90 minutes of post-addition stirring, the reaction mixture was distributed among 6 Teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and time periods. The products were isolated by centrifugation, washed with de-ionized water, and dried at room temperature. All of the products were identified as SAPO-67 based on their x-ray diffraction patterns. The characteristic diffraction lines for the product isolated from the 175°C/38 hour digestion are shown in Table 6 below.

Table 6

2- $\Theta$	d(Å)	I/I <sub>0</sub> %
8.58	10.30	w
10.90	8.11	m
11.60	7.62	w
13.28	6.66	m
15.84	5.59	w
17.24	5.14	s
17.68	5.01	w
20.94	4.24	m
21.82	4.07	vs
22.62	3.93	w
23.15	3.84	m
23.38	3.80	w
24.90	3.57	w
26.80	3.32	m
28.48	3.13	m
29.02	3.07	w
31.56	2.83	w
32.04	2.79	m
34.38	2.61	w
34.86	2.57	w
35.42	2.53	w
35.71	2.51	w
36.91	2.43	w
37.66	2.39	w
39.27	2.29	w
40.69	2.22	w
41.89	2.15	w
42.28	2.14	w



42.78	2.11	w
47.90	1.90	w
48.40	1.88	w
49.42	1.84	w
51.20	1.78	w
55.03	1.67	w
55.25	1.66	w

## EXAMPLE 7

[0042] A Teflon beaker was charged with 130.00 g ETMAOH (20 %) and placed under a high speed stirrer. With vigorous stirring, 6.29 g Al(OH)<sub>3</sub> (85.1 %) was added slowly, forming a cloudy suspension. This was followed by the dropwise addition of 33.25 g H<sub>3</sub>PO<sub>4</sub> (85 %) and further homogenization that yielded a white homogenous reaction mixture. Separately, 2.94 g Mg(OAc)<sub>2</sub>\*4H<sub>2</sub>O was dissolved in 10.30 g de-ionized water. This solution was added dropwise. Near the end of the addition the reaction mixture became a thick gel. The gel thinned out with some very high speed stirring and the resulting smooth white gel was homogenized further with normal stirring. The reaction mixture was distributed among 8 Teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and time periods. The reaction products were isolated by centrifugation, washed with de-ionized water and dried at room temperature. The reaction mixtures digested at 125 and 150 °C yielded products identified as MAPO-67 by powder x-ray diffraction. Characteristic diffraction lines for the product from a 125°C/155 hour digestion are given below in Table 7. A portion of this material was calcined at 500°C, using a ramp rate of 2°C/min in a nitrogen atmosphere, calcining in nitrogen for 4 hours before switching to dry air and calcining for an additional 2 hours at 500°C. The material was identified as AIPO-67C by powder x-ray diffraction. The diffraction lines for the calcined material are shown in table 7 below.

Table 7

Example 7			Example 7 calcined		
2- $\Theta$	d(Å)	I/I <sub>0</sub> %	2- $\Theta$	d(Å)	I/I <sub>0</sub> %
8.65	10.22	w	8.68	10.18	8.68
10.91	8.11	m	10.96	8.07	10.96

11.55	7.66	w	11.56	7.65	11.56
13.42	6.59	m	13.5	6.55	13.5
15.98	5.54	w	16.11	5.5	16.11
17.33	5.11	s	17.4	5.09	17.4
17.75	4.99	w	20.9	4.25	20.9
20.82	4.26	m	22.14	4.01	22.14
21.98	4.04	vs	23.28	3.82	23.28
22.74	3.91	w	23.59	3.77	23.59
23.36	3.81	m	24.98	3.56	24.98
24.88	3.58	w	27.16	3.28	27.16
27.02	3.30	w	28.38	3.14	28.38
28.34	3.15	m	31.57	2.83	31.57
29.25	3.05	w	32.46	2.76	32.46
31.44	2.84	w	34.47	2.6	34.47
32.30	2.77	m			
34.34	2.61	w			
35.13	2.55	w			
36.04	2.49	w			
39.65	2.27	w			
41.94	2.15	w			
42.64	2.12	w			
44.88	2.02	w			
47.55	1.91	w			
48.53	1.87	w			
49.16	1.85	w			
49.85	1.83	w			
51.18	1.78	w			
55.10	1.67	w			
55.77	1.65	w			

## EXAMPLE 8

[0043] A Teflon bottle was charged with 130.00 g ETMAOH (20 %) and 6.19 g LUDOX AS-40<sup>TM</sup> colloidal silica (40 % SiO<sub>2</sub>), stirred briefly, sealed and placed in a 95°C oven for 2 hours. The resulting solution was placed in a Teflon beaker under a high speed stirrer. Began adding 6.29 g Al(OH)<sub>3</sub> (85.1 %) slowly with vigorous stirring. The resulting cloudy suspension was homogenized for an additional 45 minutes one the addition was complete. Then 28.50 g H<sub>3</sub>PO<sub>4</sub> (85%) was added dropwise and intermittently in 5 separate portions, resulting in a semi-transparent homogeneous reaction mixture. Separately, 3.01 g Zn(OAc)<sub>2</sub>\*2H<sub>2</sub>O was dissolved in 12.0 g de-ionized water. This solution was added dropwise intermittently in three separate portions. The opaque reaction mixture was distributed among 7 Teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and time periods. The reaction products were isolated by centrifugation, washed with de-ionized water and dried at room temperature. The products from the 125, 150 and 175 °C digestions were identified as MAPSO-67 via powder x-ray diffraction. Characteristic diffraction lines for the product from the 175°C/182 hour digestion are given in Table 8 below.

Table 8

2- $\Theta$	d(Å)	I/I <sub>0</sub> %	
8.52	10.37	w	br
10.67	8.29	m	br
13.32	6.64	vs	
17.26	5.13	m	br
21.00	4.23	m	sh
21.82	4.07	s	br
23.18	3.83	s	
26.82	3.32	m	
28.49	3.13	m	br
28.88	3.09	m	sh
32.06	2.79	m	br
49.44	1.84	w	
50.94	1.79	m	
br = broad, sh = shoulder			

## EXAMPLE 9

[0044] A Teflon bottle was charged with 130.00 g ETMAOH (20 %) and 6.19 g LUDOX AS-40<sup>TM</sup> colloidal silica (40 % SiO<sub>2</sub>), stirred briefly, sealed and placed in a 95°C oven for 90 minutes. The resulting solution was placed in a Teflon beaker under a high speed stirrer.

- 5 Began adding 6.29 g Al(OH)<sub>3</sub> (85.1 %) slowly and intermittently in 5 separate portions. The resulting cloudy suspension was homogenized. Then 28.50 g H<sub>3</sub>PO<sub>4</sub> (85%) was added dropwise and intermittently, resulting in a homogeneous opalescent reaction mixture. Separately, 1.51 g Zn(OAc)<sub>2</sub>\*2H<sub>2</sub>O was dissolved in 6.00 g de-ionized water. This solution was added dropwise intermittently in three separate portions. The opaque reaction mixture
- 10 was distributed among 7 Teflon-lined autoclaves and digested at autogenous pressure at a variety of temperatures and time periods. The reaction products were isolated by centrifugation, washed with de-ionized water and dried at room temperature. The products from the 125, 150 and 175 °C digestions were identified as MAPSO-67 via powder x-ray diffraction. Characteristic diffraction lines for the product from the 175°C/42 hour digestion
- 15 are given in Table 9 below.

Table 9

2- $\Theta$	d(Å)	I/I <sub>0</sub> %	
8.52	10.37	w	br
10.92	8.10	m	br
13.36	6.62	vs	
15.84	5.59	w	sh
17.26	5.13	s	
20.88	4.25	m	sh
21.96	4.04	vs	br
23.26	3.82	s	
26.86	3.32	m	
28.48	3.13	m	br
28.96	3.08	m	sh
32.22	2.78	m	
34.50	2.60	w	
35.82	2.50	w	
42.51	2.12	w	
51.01	1.79	m	
55.56	1.65	w	
br = broad, sh = shoulder			

## EXAMPLE 10

[0045] A Teflon bottle was charged with 150.00 g DEDMAOH (20 %) and 7.56 g LUDOX AS-40<sup>TM</sup> colloidal silica (40 % SiO<sub>2</sub>), stirred briefly, sealed and placed in a 95°C oven for 2 hours. The resulting solution was placed in a Teflon beaker under a high speed stirrer. Began adding 41.83 g Al(O*sec*Bu)<sub>3</sub> (10.9 % Al) slowly resulting in a clear solution immediately. Then 23.98 g H<sub>3</sub>PO<sub>4</sub> (85.7 %) was added dropwise and intermittently, resulting in a clear solution with some separation of *sec*-BuOH formed during the hydrolysis. The reaction mixture was allowed to homogenize further. The reaction mixture was then distributed among two Teflon-lined autoclaves and digested at autogenous pressure at 175°C for 61 hours. The reaction products were isolated by centrifugation, washed with de-ionized water and dried at room temperature. The products were identified as SAPO-67 via powder x-ray diffraction. Characteristic diffraction lines for the product are given in Table 10 below. A portion of this material was calcined at 500°C first in nitrogen and then in air. The sample was ramped to 500°C at 2°C/min in nitrogen and then calcined for 4 hours in nitrogen. While the temperature remained at 500°C, the atmosphere was switched to air and the material was calcined for another two hours. The material was identified as AlPO-67C via powder x-ray diffraction. The characteristic diffraction lines for the product are shown in Table 10 below.

Table 10

Example 10			Example 10 calcined		
2- $\Theta$	d(Å)	I/I <sub>0</sub> %	2 $\Theta$	d(Å)	I/I <sub>0</sub> %
8.58	10.30	w	8.82	10.01	w
10.84	8.16	m	11.00	8.04	vs
13.32	6.64	m	13.60	6.51	vs
15.86	5.58	w	17.38	5.1	m
17.20	5.15	s	21.04	4.22	m
17.80	4.98	w	22.1	4.02	s
21.10	4.21	m	23.53	3.78	m

21.80	4.07	vs	25.13	3.54	w
23.14	3.84	m	27.12	3.29	m
25.01	3.56	w	28.58	3.12	m
26.78	3.33	m	29.36	3.04	w
28.52	3.13	m	31.64	2.83	w
28.94	3.08	w	32.30	2.77	m
30.98	2.88	w	34.54	2.6	w
31.62	2.83	w	35.28	2.54	w
31.98	2.80	m	42.72	2.11	w
34.42	2.60	w	51.28	1.78	w
34.74	2.58	w			
35.40	2.53	w			
35.66	2.52	w			
40.64	2.22	w			
41.90	2.15	w			
42.70	2.12	w			
44.17	2.05	w			
44.42	2.04	w			
48.33	1.88	w			
48.66	1.87	w			
49.34	1.85	w			
51.10	1.79	w			
55.18	1.66	w			

**[0046]** One embodiment involves a microporous crystalline metallophosphate material having a three-dimensional framework of  $\text{EO}_2^-$  and  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and

SiO<sub>2</sub> tetrahedral units and an empirical composition in the as synthesized and anhydrous basis expressed by an empirical formula of:



where M is at least one framework cation of +2 valence selected from the group consisting of  
 5 alkaline earth and transition metals, “m” is the mole ratio of M to E and varies from 0 to 1.0, R is an organoammonium cation selected from the group consisting of ethyltrimethylammonium (ETMA<sup>+</sup>), diethyldimethylammonium (DEDMA<sup>+</sup>), and mixtures thereof, “r” is the mole ratio of R to E and has a value of 0.1 to 2, E is a trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof,  
 10 “x” is the mole ratio of P to E and has a value of 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y” ≥ 0.02, and “z” is the mole ratio of O to (Al + E) and has a value determined by the equation:

$$z = (2 \cdot m + r + 3 + 5 \cdot x + 4 \cdot y) / 2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings  
 15 and intensities set forth in Table A1 above. M may be selected from the group consisting of beryllium, magnesium, zinc, cobalt, manganese, iron(II), nickel and mixtures thereof. E may be aluminum. “y” may be zero. “m” may be zero. “m” and “y” may be each greater than zero and “m” + “y” ≥ 0.02. The metallophosphate material may be thermally stable up to a temperature of at least 400°C. The metallophosphate material may be thermally stable up to a  
 20 temperature of at least 500°C. R may be ETMA<sup>+</sup> or may be DEDMA<sup>+</sup>, or may be a mixture of ETMA<sup>+</sup> and DEDMA<sup>+</sup>.

**[0047]** One embodiment involves process for preparing a microporous crystalline metallophosphate having a three-dimensional framework of EO<sub>2</sub><sup>-</sup>, PO<sub>2</sub><sup>+</sup> and at least one of [M<sup>2+</sup>O<sub>2</sub>]<sup>2-</sup> and SiO<sub>2</sub> tetrahedral units and a composition in the as synthesized and anhydrous  
 25 basis expressed by an empirical formula of:

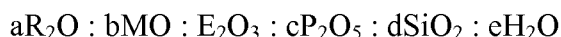


where M is at least one framework cation of valence 2+ selected from the group consisting of alkaline earth and transition metals, “m” is the mole ratio of M to E and varies from 0 to 1.0, R is an organoammonium cation selected from the group consisting of  
 30 ethyltrimethylammonium (ETMA<sup>+</sup>), diethyldimethylammonium (DEDMA<sup>+</sup>), and mixtures

thereof, “r” is the mole ratio of R to E and has a value of 0.1 to 2.0, E is a trivalent element selected from the group consisting of aluminum, gallium, iron(III), boron and mixtures thereof, “x” is the mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (2 \cdot m + r + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1 above and where the process comprises forming a reaction mixture containing reactive sources of R, E, P, at least one of M and Si, and heating the reaction mixture at a temperature of 60°C to 200°C, for a time sufficient to form the metallophosphate, the reaction mixture having a composition expressed in terms of mole ratios of the oxides of:

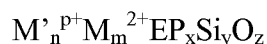


where “a” has a value of 0.25 to 16, “b” has a value of 0 to 2.0, “c” has a value of 0.8 to 8, “d” has a value of 0 to 4, and “e” has a value of 30 to 800 and wherein the sum of the values of “b” and “d” is greater than 0.01. M may be selected from the group consisting of beryllium, magnesium, zinc, cobalt, manganese, nickel, iron (II) and mixtures thereof. The source of M may be selected from the group consisting of halide salts, nitrate salts, acetate salts, sulfate salts and mixtures thereof. The source of E may be selected from the group consisting of aluminum isopropoxide, aluminum sec-butoxide, precipitated alumina, Al(OH)<sub>3</sub>, aluminum metal, aluminum salts, boric acid, precipitated gallium oxyhydroxide, gallium sulfate, ferric sulfate, ferric chloride and mixtures thereof. The silicon source may be selected from the group consisting of tetraethyorthosilicate, fumed silica, colloidal silica and precipitated silica. The reaction mixture may be reacted at a temperature of 125°C to 175°C for a time of 1 day to 10 days. R may be ETMA<sup>+</sup>. R may be DEDMA<sup>+</sup>. R may be a combination of ETMA<sup>+</sup> and DEDMA<sup>+</sup>. AIPO-67 seeds may be added to the reaction mixture.

**[0048]** Another embodiment is a modified form of the above metallophosphate where the modification comprises ammonia calcination under ammonia calcinations conditions, and optionally an additional ion-exchange to form a modified microporous metallophosphate,



AlPO-67M, having a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units the composition given by the empirical formula:



where M is at least one metal cation of valence +2 selected from the group consisting of  $\text{Be}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Ni}^{2+}$ , “m” is the mole ratio of M to E and varies from 0 to 1.0, M’ is selected from the group consisting of  $\text{NH}_4^+$ ,  $\text{H}^+$ , alkali metals, alkaline earth metals, rare earth metals and mixtures thereof, “n” is the mole ratio of M’ to E and has a value of 0.03 to 2.0, “p” is the weighted average valence of M’ and varies from 1 to 3, E is an trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (p \cdot n + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1 above.

**[0049]** Another embodiment is a stable calcined form of the crystalline microporous metallophosphate, AlPO-67C, comprising heating the metallophosphate under calcination conditions to remove the organic cations in the presence of air and ambient water vapor, said AlPO-67C comprising a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units characterized by the empirical formula:



where M is at least one metal cation of valence +2 selected from the group consisting of  $\text{Be}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Ni}^{2+}$ , “m” is the mole ratio of M to E and varies from 0 to 1.0, H is a proton, “a” is the mole ratio of H to E and has a value of 0.1 to 2.0, E is an trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (a + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table B above

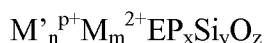
[0050] Another embodiment involves a hydrocarbon conversion process comprising contacting a stream comprising at least one hydrocarbon with a catalyst at hydrocarbon conversion conditions to generate at least one converted product, wherein the catalyst is selected from the group consisting of a microporous crystalline AIPO-67 material, a microporous crystalline AIPO-67M material, a microporous crystalline AIPO-67-C material, or mixtures thereof, where AIPO-67 has a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units and an empirical composition in the as synthesized and anhydrous basis expressed by an empirical formula of:



where M is at least one framework cation of valence 2+ selected from the group consisting of alkaline earth and transition metals, “m” is the mole ratio of M to E and varies from 0 to 1.0, R is an organoammonium cation selected from the group of ethyltrimethylammonium (ETMA<sup>+</sup>), diethyldimethylammonium (DEDMA<sup>+</sup>), and mixtures thereof, “r” is the mole ratio of R to E and has a value of 0.1 to 2.0, E is a trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is the mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y” ≥ 0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (2 \cdot m + r + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1 above where the microporous crystalline AIPO-67M material comprises a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units the composition given by the empirical formula:

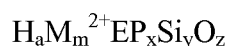


where M is at least one metal cation of valence +2 selected from the group consisting of  $\text{Be}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Ni}^{2+}$ , “m” is the mole ratio of M to E and varies from 0 to 1.0, M' is selected from the group consisting of  $\text{NH}_4^+$ ,  $\text{H}^+$ , alkali metals, alkaline earth metals, rare earth metals, and mixtures thereof, “n” is the mole ratio of M' to E and has a value of 0.03 to 2.0, “p” is the weighted average valence of M' and varies from 1 to 3, E is an trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole

ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (p \cdot n + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

5 and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1 above and the microporous crystalline AlPO-67C material comprises a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$ ,  $\text{SiO}_2$  and optionally  $[\text{M}^{2+}\text{O}_2]^{2-}$  tetrahedral units characterized by the empirical formula:



10 where M is at least one metal cation of valence +2 selected from the group consisting of  $\text{Be}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Ni}^{2+}$ , “m” is the mole ratio of M to E and varies from 0 to 1.0, H is a proton, “a” is the mole ratio of H to E and has a value of 0.1 to 2.0, E is an trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E  
15 and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (a + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table B above. The hydrocarbon conversion process may be  
20 selected from the group consisting of cracking, hydrocracking, alkylation, isomerization, polymerization, reforming, hydrogenation, dehydrogenation, transalkylation, dealkylation, hydration, dehydration, hydrotreating, hydrofining, hydrodenitrogenation, hydrodesulfurization, methanol to olefins, methanation, syngas shift process, olefin dimerization, oligomerization, dewaxing, and combinations thereof. The hydrocarbon  
25 conversion process may be hydrocracking or hydrotreating and wherein the hydrocracking or hydrotreating is operated at a temperature in the range of 400° to 1200°F (204-649°C) and a pressure in the range of atmospheric to 3,500 psig (24,132 kPa g). The process may be hydrocracking or hydrotreating operated at a liquid hourly space velocities (LHSV) in the range of 0.1  $\text{hr}^{-1}$  to 15  $\text{hr}^{-1}$ . The process may be hydrocracking or hydrotreating operated at

hydrogen circulation rates in the range of 1,000 to 50,000 standard cubic feet (scf) per barrel of charge (178-8,888 std. m<sup>3</sup>/m<sup>3</sup>). The process may further comprise removing an effluent comprising the at least one converted product, fractionating the effluent, and recovering at least one converted product. The process may further comprise, subjecting the effluent to partial condensation and vapor-liquid separation prior to fractionation. The process may further comprise recycling at least a portion of the effluent to the catalyst. The process may comprise two stage operation and the catalyst is present in at least one of the two stages. The hydrocarbon conversion process may be catalytic cracking operated at a temperature in the range of 850° to 1100°F, LHSV values of 0.5 to 10 and a pressure in the range of from 0 to 50 psig. The stream comprising at least one hydrocarbon may be selected from the group consisting of gas oils, heavy naphthas, and deasphalted crude oil residua. The hydrocarbon conversion process may be alkylation of aromatics and the converted product may be at least one linear alkyl substituted aromatic when the process is operated at an aromatic: olefin mole ratio of between 5:1 and 30:1, a LHSV of 0.3 to 6 hr<sup>-1</sup>, a temperature of 100° to 250°C and a pressures of 200 to 1000 psig. The hydrocarbon conversion process may be alkylation of isoparaffins with olefins and the converted product may be at least one alkylate suitable as a motor fuel component, and the process may be operated at a temperature of from -30° to 40°C, a pressure from atmospheric to 6,894 kPa (1,000 psig) and a weight hourly space velocity (WHSV) of 0.1 to 120. The hydrocarbon conversion process may be the conversion of methanol to olefins wherein the process is operated at a temperature of 300 °C to 600 °C and a pressure from 0 kPa (0 psig) to 1724 kPa (250 psig). The methanol may be in the liquid or vapor phase and the operation maybe in a continuous or batch mode. The methanol maybe diluted with a diluent selected from the group consisting of helium, argon, nitrogen, carbon monoxide, carbon dioxide, hydrogen, steam, at least one paraffinic hydrocarbon, at least one aromatic hydrocarbon, and mixtures thereof. The catalyst may be located in one or more catalyst zones arranged in series or parallel configuration, and wherein the catalyst may be in fixed beds or fluidized beds. The process may further comprise regenerating the catalyst.

**[0051]** Another embodiment involves a separation process comprising contacting a mixture of at least two components with a material to generate at least one separated component, wherein the material is selected from the group consisting of a microporous crystalline AlPO-67 material, a microporous crystalline AlPO-67M material, a microporous

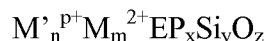
crystalline AIPO-67-C material, or mixtures thereof, where AIPO-67 has a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units and an empirical composition in the as synthesized and anhydrous basis expressed by an empirical formula of:



where M is at least one framework cation of valence 2+ selected from the group consisting of alkaline earth and transition metals, “m” is the mole ratio of M to E and varies from 0 to 1.0, R is an organoammonium cation selected from the group of ethyltrimethylammonium ( $\text{ETMA}^+$ ), diethyldimethylammonium ( $\text{DEDMA}^+$ ), and mixtures thereof, “r” is the mole ratio of R to E and has a value of 0.1 to 2.0, E is a trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is the mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (2 \cdot m + r + 3 + 5 \cdot x + 4 \cdot y)/2$$

15 and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1 above and the microporous crystalline AIPO-67M material comprises a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units the composition given by the empirical formula:



where M is at least one metal cation of valence +2 selected from the group consisting of  $\text{Be}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Ni}^{2+}$ , “m” is the mole ratio of M to E and varies from 0 to 1.0, M' is selected from the group consisting of  $\text{NH}_4^+$ ,  $\text{H}^+$ , alkali metals, alkaline earth metals, rare earth metals and mixtures thereof, “n” is the mole ratio of M' to E and has a value of 0.03 to 2.0, “p” is the weighted average valence of M' and varies from 1 to 3, E is an trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (p \cdot n + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1 above and the microporous crystalline AlPO-67C material comprises a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units characterized by the empirical formula:



where M is at least one metal cation of valence +2 selected from the group consisting of  $\text{Be}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Ni}^{2+}$ , “m” is the mole ratio of M to E and varies from 0 to 1.0, H is a proton, “a” is the mole ratio of H to E and has a value of 0.1 to 2.0, E is an trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, “x” is mole ratio of P to E and varies from 0.5 to 2.0, “y” is the mole ratio of Si to E and varies from 0 to 1.0, “m” + “y”  $\geq$  0.02, and “z” is the mole ratio of O to E and has a value determined by the equation:

$$z = (a + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table B above. The separation may be based on molecular size of the components, degree of polarity of the components, or ion exchange of the components with the material.

## WHAT IS CLAIMED IS:

1. A microporous crystalline metallophosphate material having a three-dimensional framework of  $\text{EO}_2^-$  and  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units and an empirical composition in the as synthesized and anhydrous basis expressed by an empirical formula of:



where M is at least one framework cation of +2 valence selected from the group consisting of alkaline earth and transition metals, "m" is the mole ratio of M to E and varies from 0 to 1.0, R is an organoammonium cation selected from the group consisting of ethyltrimethylammonium ( $\text{ETMA}^+$ ), diethyldimethylammonium ( $\text{DEDMA}^+$ ), and mixtures thereof, "r" is the mole ratio of R to E and has a value of 0.1 to 2, E is a trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, "x" is the mole ratio of P to E and has a value of 0.5 to 2.0, "y" is the mole ratio of Si to E and varies from 0 to 1.0, "m" + "y"  $\geq$  0.02, and "z" is the mole ratio of O to (Al + E) and has a value determined by the equation:

$$z = (2 \cdot m + r + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1:

Table A1

$2\Theta$	$d(\text{\AA})$	$I/I_0\%$
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s
13.65 - 13.16	6.48 - 6.72	m - vs
17.72 - 16.91	5.00 - 5.24	m - vs
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.5	3.96 - 4.13	s - vs
23.84 - 22.84	3.73 - 3.89	w - vs
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m
36.65 - 35.31	2.45 - 2.54	w - m

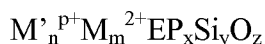
52.23 - 50.67	1.75 - 1.80	w - m
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2. The metallophosphate material of claim 1 where M is selected from the group consisting of beryllium, magnesium, zinc, cobalt, manganese, iron(II), nickel and mixtures thereof.

5 3. The metallophosphate material of claim 1 where the metallophosphate material is thermally stable up to a temperature of at least 400°C.

4. The metallophosphate material of claim 1 where R is ETMA<sup>+</sup> or DEDMA<sup>+</sup>, or a mixture of ETMA<sup>+</sup> and DEDMA<sup>+</sup>.

10 5. The metallophosphate material of claim 1, after modification comprising ammonia calcination under ammonia calcinations conditions, and optionally an additional ion-exchange to form a modified microporous metallophosphate, AlPO-67M, having a three-dimensional framework of EO<sub>2</sub><sup>-</sup>, PO<sub>2</sub><sup>+</sup> and at least one of [M<sup>2+</sup>O<sub>2</sub>]<sup>2-</sup> and SiO<sub>2</sub> tetrahedral units the composition given by the empirical formula:



15 where M is at least one metal cation of valence +2 selected from the group consisting of Be<sup>2+</sup>, Mg<sup>2+</sup>, Zn<sup>2+</sup>, Co<sup>2+</sup>, Mn<sup>2+</sup>, Fe<sup>2+</sup>, Ni<sup>2+</sup>, "m" is the mole ratio of M to E and varies from 0 to 1.0, M' is selected from the group consisting of NH<sub>4</sub><sup>+</sup>, H<sup>+</sup>, alkali metals, alkaline earth metals, rare earth metals and mixtures thereof, "n" is the mole ratio of M' to E and has a value of 0.03 to 2.0, "p" is the weighted average valence of M' and varies from 1 to 3, E is an trivalent element selected  
20 from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, "x" is mole ratio of P to E and varies from 0.5 to 2.0, "y" is the mole ratio of Si to E and varies from 0 to 1.0, "m" + "y" ≥ 0.02, and "z" is the mole ratio of O to E and has a value determined by the equation:

$$z = (p \cdot n + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

25 and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1:

Table A1

2θ	d(Å)	I/I <sub>0</sub> %
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s



13.65 - 13.16	6.48 - 6.72	m - vs
17.72 - 16.91	5.00 - 5.24	m - vs
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.5	3.96 - 4.13	s - vs
23.84 - 22.84	3.73 - 3.89	w - vs
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m
36.65 - 35.31	2.45 - 2.54	w - m
52.23 - 50.67	1.75 - 1.80	w - m

6. The metallophosphate material of claim 1, after calcination comprising heating the metallophosphate of claim 1 under calcination conditions to remove the organic cations in the presence of air and ambient water vapor, to form AlPO-67C comprising a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units characterized by the empirical formula:



where M is at least one metal cation of valence +2 selected from the group consisting of  $\text{Be}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$ ,  $\text{Ni}^{2+}$ , "m" is the mole ratio of M to E and varies from 0 to 1.0, H is a proton, "a" is the mole ratio of H to E and has a value of 0.1 to 2.0, E is a trivalent element selected from the group consisting of aluminum, gallium, iron, boron and mixtures thereof, "x" is mole ratio of P to E and varies from 0.5 to 2.0, "y" is the mole ratio of Si to E and varies from 0 to 1.0, "m" + "y"  $\geq$  0.02, and "z" is the mole ratio of O to E and has a value determined by the equation:

$$z = (a + 2 \cdot m + 3 + 5 \cdot x + 4 \cdot y)/2$$

and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table B:

Table B

2 $\Theta$	d( $\text{\AA}$ )	I/I <sub>0</sub> %
9.02 - 8.33	9.80 - 10.6	w - m
11.09 - 10.79	7.97 - 8.19	vs
13.74 - 13.32	6.44 - 6.64	s - vs
17.58 - 16.97	5.04 - 5.22	m
21.19 - 20.69	4.19 - 4.29	m

22.32 - 21.87	3.98 - 4.06	m - s
28.78 - 28.13	3.10 - 3.17	m
32.78 - 32.05	2.73 - 2.79	w - m

7. A process for preparing a microporous crystalline metallophosphate having a three-dimensional framework of  $\text{EO}_2^-$ ,  $\text{PO}_2^+$  and at least one of  $[\text{M}^{2+}\text{O}_2]^{2-}$  and  $\text{SiO}_2$  tetrahedral units and a composition in the as synthesized and anhydrous basis expressed by an empirical formula of:



where M is at least one framework cation of valence 2+ selected from the group consisting of alkaline earth and transition metals, "m" is the mole ratio of M to E and varies from 0 to 1.0, R is an organoammonium cation selected from the group consisting of ethyltrimethylammonium ( $\text{ETMA}^+$ ), diethyldimethylammonium ( $\text{DEDMA}^+$ ), and mixtures thereof, "r" is the mole ratio of R to E and has a value of 0.1 to 2.0, E is a trivalent element selected from the group consisting of aluminum, gallium, iron(III), boron and mixtures thereof, "x" is the mole ratio of P to E and varies from 0.5 to 2.0, "y" is the mole ratio of Si to E and varies from 0 to 1.0, "m" + "y"  $\geq$  0.02, and "z" is the mole ratio of O to E and has a value determined by the equation:

$$z = (2 \cdot m + r + 3 + 5 \cdot x + 4 \cdot y)/2$$

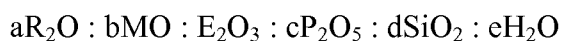
and is characterized in that it has the x-ray diffraction pattern having at least the d-spacings and intensities set forth in Table A1:

Table A1

$2\Theta$	$d(\text{\AA})$	$I/I_0\%$
8.86 - 8.39	9.97 - 10.53	w - m
11.21 - 10.51	7.89 - 8.41	m - s
13.65 - 13.16	6.48 - 6.72	m - vs
17.72 - 16.91	5.00 - 5.24	m - vs
21.29 - 20.69	4.17 - 4.29	m - s
22.43 - 21.5	3.96 - 4.13	s - vs
23.84 - 22.84	3.73 - 3.89	w - vs
27.51 - 26.51	3.24 - 3.36	w - m
29.06 - 27.95	3.07 - 3.19	w - m
32.90 - 31.59	2.72 - 2.83	m
35.02 - 33.93	2.56 - 2.64	w - m

36.65 - 35.31	2.45 - 2.54	w - m
52.23 - 50.67	1.75 - 1.80	w - m

the process comprising forming a reaction mixture containing reactive sources of R, E, P, at least one of M and Si, and heating the reaction mixture at a temperature of 60°C to 200°C, for a time sufficient to form the metallophosphate, the reaction mixture having a composition  
 5 expressed in terms of mole ratios of the oxides of:



where “a” has a value of 0.25 to 16, “b” has a value of 0 to 2.0, “c” has a value of 0.8 to 8, “d” has a value of 0 to 4, and “e” has a value of 30 to 800 and wherein the sum of the values of “b” and “d” is greater than 0.01.

10           8.       A process of contacting a feed with any of the microporous crystalline metallophosphates of claims 1, 8, or 9 wherein the process is a hydrocarbon conversion process or a separation process.

15           9.       The process of claim 7 wherein the hydrocarbon conversion process is selected from the group consisting of cracking, hydrocracking, alkylation, isomerization, polymerization, reforming, hydrogenation, dehydrogenation, transalkylation, dealkylation, hydration, dehydration, hydrotreating, hydrofining, hydrodenitrogenation, hydrodesulfurization, methanol to olefins, methanation, syngas shift process, olefin dimerization, oligomerization, dewaxing, and combinations thereof.

20           10.       The process of claim 7 wherein the separation process is based on molecular size of the components, degree of polarity of the components, or ion exchange of the components with the material.

**INTERNATIONAL SEARCH REPORT**

International application No.

PCT/US 2013/071562

A. CLASSIFICATION OF SUBJECT MATTER  
(see extra sheet)

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)  
B01J 27/14, 31/02, 27/18, 27/182, 29/00, 29/82, C01B 25/26, 25/37, 39/46, 39/54, C07C 1/20, 5/22, 2/66, 5/27, C07F 11/00, 15/00, 15/02, C10G 11/02, 35/06, 47/20,

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)  
PatSearch (RUPTO internal), USPTO, PAJ, Esp@cenet, DWPI, EAPATIS, PATENTSCOPE, Information Retrieval System of FIPS

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 4973785 A (UOP) 27.11.1990	1-10
A	US 4440871 A (UNION CARBIDE CORPORATION) 03.04.1984	1-10
A	US 4567029 A (UNION CARBIDE CORPORATION) 28.01.1986	1-10
A	US 4310440 A (UNION CARBIDE CORPORATION) 12.01.1982	1-10
A	US 5200187 A (EXXOM RESEARCH AND ENGINEERING COMPANY) 06.04.1993	1-10
A	US 2012/0157741 A1 (GUANG CAO et al.) 21.06.2012	1-10
A	US 6613302 B1 (UOP LLC) 02.09.2003	1-10

Further documents are listed in the continuation of Box C.

See patent family annex.

* Special categories of cited documents:	“T” later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
“A” document defining the general state of the art which is not considered to be of particular relevance	“X” document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
“E” earlier document but published on or after the international filing date	“Y” document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
“L” document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)	“&” document member of the same patent family
“O” document referring to an oral disclosure, use, exhibition or other means	
“P” document published prior to the international filing date but later than the priority date claimed	

Date of the actual completion of the international search 14 February 2014 (14.02.2014)	Date of mailing of the international search report 17 April 2014 (17.04.2014)
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**INTERNATIONAL SEARCH REPORT**

International application No.

PCT/US 2013/071562

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 6388159 B1 (UOP LLC) 14.05.2002	1-10
A	WO 2005/113439 A1 (UOP LLC et al.) 01.12.2005	1-10

***B01J 27/14 (2006.01)***  
***B01J 31/02 (2006.01)***  
***B01J 27/18 (2006.01)***  
***B01J 27/182 (2006.01)***  
***B01J 29/82 (2006.01)***  
***C01B 25/26 (2006.01)***  
***C01B 39/54 (2006.01)***  
***C07C 5/22 (2006.01)***  
***C07C 2/66 (2006.01)***  
***C07F 15/00 (2006.01)***  
***C07F 15/02 (2006.01)***  
***C10G 11/02 (2006.01)***  
***C10G 35/06 (2006.01)***  
***C10G 47/20 (2006.01)***  
***C01B 39/46 (2006.01)***