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Iron complexes

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IRON COMPLEXES

This invention relates to iron compounds for use in pharmaceutical compositions for the treatment of iron deficiency anaemia.

05 An adequate supply of iron to the body is an essential requirement for tissue growth in both man and animals. Although there is normally an ample amount of iron in the diet, the level of absorption of iron from food is generally low so that the supply of iron to the body can easily become critical under a variety of conditions. Iron deficiency anaemia is
10 commonly encountered in pregnancy and may also present a problem in the newly born, particularly in certain animal species such as the pig. Moreover, in certain pathological conditions there is a mal distribution of body iron leading to a state of chronic anaemia. This is seen in chronic diseases
15 such as rheumatoid arthritis, certain haemolytic diseases and cancer.

Although a wide range of iron compounds is already marketed for the treatment of iron deficiency anaemia, the level of iron uptake by the body from these compounds is often quite low,
20 necessitating the administration of relatively high dosage levels of the compound. The administration of high dose, poorly absorbed, iron complexes may cause siderosis of the gut wall and a variety of side effects such as nausea, vomiting, constipation and heavy malodorous stools.

25 In the UK patent applications of numbers 8308053 (published as GB 2117766A), 8327612 (published as GB 2128998A) and 8407180 (published as GB 2136806A), we describe iron complexes of various 3-hydroxypyrid-2-ones, 3-hydroxypyrid-4-ones and 3-hydroxy-4-pyrones which we have identified as being of
30 particular value for use at relatively low dosage levels in the treatment of iron deficiency anaemia. It has now been found that certain advantages, as discussed hereinafter, accrue from

the use of complexes not described in these earlier applications which contain one or more hydroxypyridone or hydroxypyrrone ligands but in which the ligands present in the complex are not identical.

- 05 According to the present invention a neutral 3:1
ligand:iron(III) complex comprises three monobasic bidentate
ligands, of which at least two are different, each ligand
separately being provided by a compound being:
- 10 (1) 3-hydroxy-4-pyrone or a 3-hydroxy-4-pyrone in which one or
more of the hydrogen atoms attached to ring carbon atoms are
replaced by an aliphatic hydrocarbon group of 1 to 6 carbon
atoms; or
- 15 (2) a 3-hydroxypyrid-2-one or 3-hydroxypyrid-4-one in which the
hydrogen atom attached to the nitrogen atom is replaced by an
aliphatic acyl group, by an aliphatic hydrocarbon group of 1
to 6 carbon atoms, or by an aliphatic hydrocarbon group
substituted by one or more substituents selected from aliphatic
acyl, alkoxy, aliphatic amide, aliphatic ester, halogen and
hydroxy groups and, optionally, in which one or more of the
20 hydrogen atoms attached to ring carbon atoms are replaced by
one of said substituents, by an aliphatic hydrocarbon group, or
by an aliphatic hydrocarbon group substituted by an alkoxy,
aliphatic ester, halogen or hydroxy group; the trivalent
positive charge on the ferric cation being neutralised by the
25 charge on the ligands covalently bound thereto.

It should be appreciated that a feature of the iron
complexes of the present invention is that they are neutral,
i.e. there is an internal balance of charges between the ferric
cation and the ligands bound covalently thereto, there being no
30 need for any additional non-covalently bound anion or anions,
such as chloride, to balance the charge on the ferric cation.
The hydroxypyridone and hydroxypyrrone ligands are each
bidentate and monobasic (the ligand containing a group $-O^-$ in
place of the group $-OH$ present in the compound itself). It is
35 preferably the case that the iron complex is a neutral 3:1

ligand:iron(III) complex which contains three monobasic, bidentate ligands, of which at least two are different, each ligand separately being provided by a compound being:

- 05 (1) 3-hydroxy-4-pyrone or a 3-hydroxy-4-pyrone as defined hereinbefore;
 - (2) a 3-hydroxypyrid-2-one or 3-hydroxypyrid-4-one as defined hereinbefore; and
 - (3) an alternative compound providing a physiologically acceptable, monobasic, bidentate ligand which is capable of
10 bonding covalently to iron;
- but with the proviso that at least one of the ligands is of type (1) or (2).

It will be appreciated that the 3:1 ligand:iron complexes of the present invention may contain various different
15 combinations of the ligands of the types (1) to (3) described above, subject to the requirement that at least one is a hydroxypyridone or hydroxypyrrone ligand and that at least two of the three ligands are not identical. Thus, all of the ligands may be of type (1) or of type (2), although differing within
20 these types, or the ligands may be of all three types. Alternatively, the ligands may be a mixture of types (1) and (2), a mixture of types (1) and (3) or a mixture of types (2) and (3). It will most usually be the case that two of the monobasic, bidentate ligands present in a 3:1 complex will be
25 the same, with the third one being different. Moreover, at least two will preferably be of the hydroxypyridone or hydroxypyrrone type. Most conveniently, therefore, the complex contains three ligands of type (2) or particularly of type (1) which are derived from two different compounds, or the complex
30 contains two identical ligands of type (2) or particularly of type (1) together with a third ligand which is either of type (1) or (2), respectively, or less preferably of type (3).

The ability of both the metal-free ligand-providing compounds and of the iron complex to permeate membranes is
35 important in the context of the treatment of iron deficiency

anaemia and it is also desirable for both to possess some degree of water solubility. A good indication of the physical properties of a ligand-providing compound and of an iron complex in this respect is provided by the value of the partition coefficient (K_{part}) obtained on partition between n-octanol and tris hydrochloride (20 mM, pH 7.4; tris representing 2-amino-2 hydroxymethylpropane 1,3-diol) at 20°C and expressed as the ratio (concentration in organic phase)/(concentration in aqueous phase). Preferred complexes show a value of K_{part} for each ligand-providing compound of above 0.02 or 0.05 but less than 3.0, especially of above 0.2 but less than 1.0, together with a value of K_{part} for the 3:1 iron(III) complex of above 0.02 but less than 6.0, especially of above 0.1 or 0.2 but less than 1.0. For examples of measured partition coefficients of metal-free hydroxypyridones and hydroxypyrones, and of iron complexes thereof in which all of the ligands are identical, reference should be made to Example 1 of each of the three applications referred to hereinbefore. The following comments upon preferences among the different ligands of types (1) to (3) which may be used in complexes according to the present invention are made in the light of these preferences as to partition coefficients.

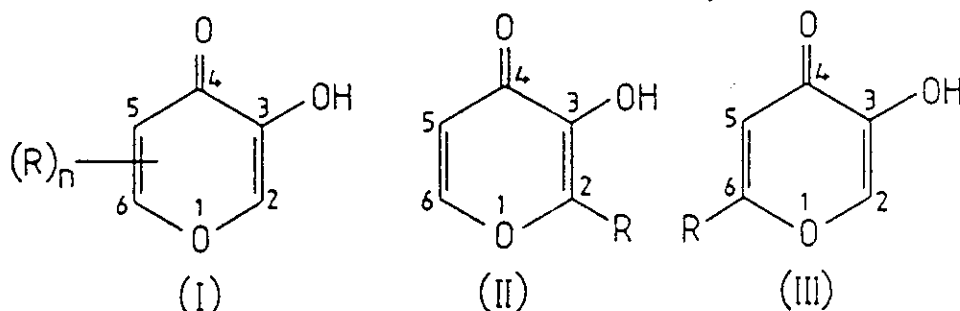
The hydroxypyrrone ligands of type (1) are of particular value and the complexes according to the present invention may conveniently contain at least one such ligand. The substituted 3-hydroxy-4-pyrones may carry more than one type of aliphatic hydrocarbon group but this is not usual and, indeed, substitution by two rather than three, and particularly by only one aliphatic hydrocarbon group is preferred. The aliphatic hydrocarbon groups may be cyclic or acyclic, having a branched chain or especially a straight chain in the latter case, and may be unsaturated or especially saturated. Groups of from 1 to 4 carbon atoms and particularly of 1 to 3 carbon atoms are of most interest. Alkyl groups are preferred, for example cyclic groups such as cyclopropyl and especially cyclohexyl but,

more particularly preferred are acyclic alkyl groups such as n-propyl and isopropyl, and especially ethyl and methyl. Substitution at the 2- or 6-position is of especial interest although, when the ring is substituted by the larger aliphatic hydrocarbon groups, there may be an advantage in avoiding substitution on a carbon atom alpha to the



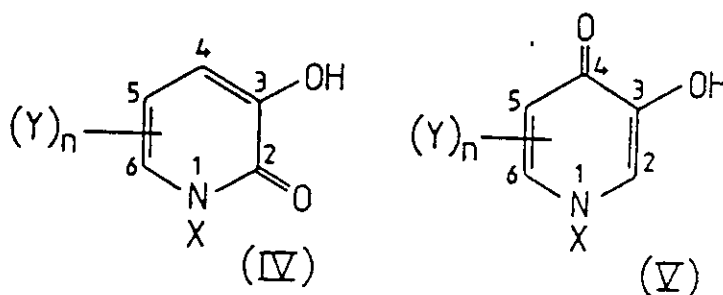
This system is involved in the complexing with iron and the close proximity of one of the larger aliphatic hydrocarbon groups may lead to steric effects which inhibit complex formation.

Preferred hydroxypyrones providing ligands present in complexes according to the present invention have the formula (I), specific hydroxypyrones of particular interest having the formulae (II) and (III):-



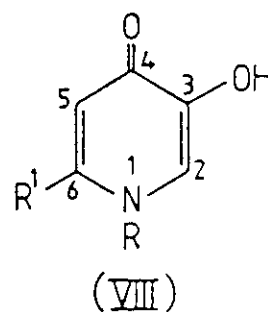
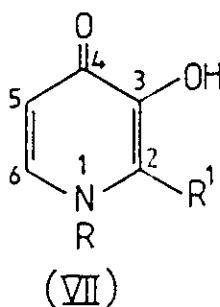
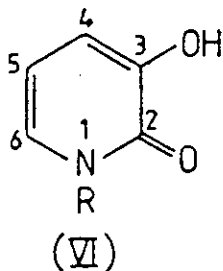
in which R is an alkyl group, for example methyl, ethyl, n-propyl isopropyl or butyl, and n is 0, 1, 2 or 3 (the ring being unsubstituted by any alkyl group when n is 0). Among these compounds 3-hydroxy-2-methyl-4-pyrone (maltol; II, R = CH₃) is of most interest, whilst 3-hydroxy-4-pyrone (pyromeconic acid; I, n = 0) 3-hydroxy-6-methyl-4-pyrone (III, R = CH₃) and particularly 2-ethyl-3-hydroxy-4-pyrone (ethylpyromeconic acid; II, R = C₂H₅) are also of especial interest.

As regards the hydroxypyridone ligands of type (2), these may be derived from hydroxypyridones of the type described in UK Patent Application No. 8308053, published as GB 2117766A, or of the type described in UK Patent Application No. 8407180, published as GB 2136806A and claiming priority from UK Patent Application No. 8308055. The former consist of a 3-hydroxypyrid-2-one or 3-hydroxypyrid-4-one in which the hydrogen atom attached to the nitrogen atom is replaced by an aliphatic hydrocarbon group of 1 to 6 carbon atoms and, optionally, in which one or more of the hydrogen atoms attached to ring carbon atoms are also replaced by the same or a different aliphatic hydrocarbon group of 1 to 6 carbon atoms, whilst the latter consist of a 3-hydroxypyrid-2-one or 3-hydroxypyrid-4-one which is substituted as defined under (2) hereinbefore but excluding those compounds in which the replacement of hydrogen atoms is effected only by aliphatic hydrocarbon groups (these compounds being the substituted hydroxypyridones of the former application). Hydroxypyridones providing ligands which may be used in complexes according to the present invention have the formulae (IV) and (V)



in which X and Y are substituents as defined hereinbefore and n is 0, 1, 2 or 3, the 3-hydroxypyrid-2-ones generally being of somewhat greater interest than the 3-hydroxypyrid-4-ones.

Preferences as to the nature and position of the substituent groups present in the hydroxypyridones are broadly as expressed in the two earlier applications. Thus, substituted aliphatic hydrocarbon groups present in the hydroxypyridones may as indicated carry more than one substituent group but it is preferred that two rather than three, and particularly only one substituent group is present. Such substituted aliphatic hydrocarbon group substituents may conveniently contain groups of 1 to 8 and particularly of 1 to 6 carbon atoms, but the simpler hydroxypyridones of UK Patent Application GB 2117766A containing only unsubstituted aliphatic hydrocarbon group substituents are of the greatest interest. The preferences among the aliphatic hydrocarbon groups present in these hydroxypyridones correspond largely to those expressed in relation to the hydroxypyrones, with methyl groups conveniently being used for substitution on ring carbon atoms but larger alkyl groups also being of particular interest for substitution on the ring nitrogen atoms. Substitution of the ring carbon atoms, which is again preferably by one rather than two or three aliphatic hydrocarbon groups, is of particular interest in the case of the 3-hydroxypyrid-4-ones, for example at the 6- or particularly the 2-position, whilst the 3-hydroxypyrid-2-ones may more often be used without any additional aliphatic hydrocarbon group substituent on the ring carbon atoms. Specific hydroxypyridones of particular interest have formulae (VI), (VII), and (VIII)



in which R is an alkyl group, for example methyl, ethyl, n-propyl, isopropyl or butyl, and R¹ is hydrogen or particularly an alkyl group, for example methyl. Such compounds include 3-hydroxy-1-methylpyrid-2-one, 3-hydroxy-1,2-dimethylpyrid-4-one and particularly 1-ethyl-3-hydroxypyrid-2-one, 3-hydroxy-1-propylpyrid-2-one, 3-hydroxy-1-(1'-methyl-ethyl)-pyrid-2-one, 1-butyl-3-hydroxypyrid-2-one, 1-ethyl-3-hydroxy-2-methylpyrid-4-one, 3-hydroxy-2-methyl-1-propylpyrid-4-one, 3-hydroxy-1-(1'-methylethyl)-2-methylpyrid-4-one and 1-butyl-3-hydroxy-2-methylpyrid-4-one with the 3-hydroxypyrid-2-ones such as 1-ethyl-3-hydroxypyrid-2-one being especially preferred.

The ligands of type (3) may be derived from various forms of compound, many of which are naturally occurring, and include physiologically acceptable, monobasic, bidentate ligands known in the art. The compounds which provide such ligands will generally comprise (a) a first grouping containing an acidic proton which is lost to provide both the single negative charge on the ligand and also one of its chelating sites and (b) a second grouping which provides the second chelating site. The grouping (a) is preferably either an enolic hydroxy group or a carboxy group whilst the grouping (b) is preferably an amino group, conveniently a primary amino group, or a hydroxy group. In a particular case, one grouping can fulfil both function (a) and function (b). Thus, some monocarboxylic acids can provide an anion capable of a bidentate mode and containing a grouping $-\text{C} \begin{smallmatrix} \text{O}^- \\ \text{O}^- \end{smallmatrix}$ which will fulfil both functions.

Apart from such monocarboxylic acids, for example formic acid, propionic acid and particularly acetic acid, many other forms of acid are of interest for providing type (3) ligands. These include various hydroxy acids, for example lactic acid, gluconic acid, etc., and various amino acids, for example glycine, isoleucine, leucine, methionine, phenylalanine, tyrosine and valine. Also of interest are peptides,

particularly the smaller compounds such as tri- and especially di-peptides, for example those containing the same or different amino acids selected from those listed above such as glycyl-leucine, leucyl-glycine and especially glycyl-glycine and leucyl-leucine. Apart from the carboxylic acids, the other group of compounds of particular interest is those containing a grouping $\text{>C}=\overset{\text{I}}{\text{C}}\text{-OH}$, a particular example of this class being ascorbic acid (vitamin C). It should be noted that ascorbic acid is capable of providing a dibasic rather than a monobasic anion but compounds such as this are quite suitable for use in providing ligands of type (3) where they have a single pK_a , only, which is less than 10 since, in use under physiological conditions, the ascorbate or other such anion will be monobasic. It will be appreciated from the foregoing discussion that the carbohydrate compounds gluconic acid and ascorbic acid are of interest in providing type (3) ligands and this interest extends to other carbohydrates, including particularly the monosaccharide sugars and related compounds. In selecting carbohydrate or other compounds for providing type (3) ligands, the more hydrophobic compounds are generally of greater interest so that among the amino acids, for example, the more complex amino acids than glycine may be of greater value.

Examples of specific iron complexes according to the present invention are (1-ethyl-3-hydroxypyrid-2-one)₂ (1-butyl-3-hydroxypyrid-4-one) iron(III), (maltol)₂ (1-ethyl-3-hydroxypyrid-2-one) iron(III), (maltol)₂ (leucine) iron(III), (maltol)₂ (glycine) iron(III), (maltol)₂ (ascorbic acid) iron(III), (maltol)₂ (gluconic acid) iron(III) and especially (maltol)₂ (ethylpyromeconic acid) iron(III) and (ethylpyromeconic acid)₂ maltol iron(III). It will be appreciated that "maltol" is used in the names of these complexes to represent the ligand derived from maltol, and

similarly for the other ligands, this usage being employed throughout the specification.

The iron complexes are conveniently prepared by the reaction of a mixture of the compounds which provide the
05 ligands and iron ions, the latter conveniently being derived from an iron salt, particularly a ferric halide and especially ferric chloride. The reaction is conveniently effected in a suitable mutual solvent and water may be used for this purpose. If desired, however, an aqueous/organic solvent
10 mixture may be used or an organic solvent, for example ethanol, methanol or chloroform and mixtures of these solvents together and/or with water where appropriate. In particular, methanol or especially ethanol may be used where it is desired to effect the separation of at least a major part of a by-product such as
15 sodium chloride by precipitation whilst the iron complex is retained in solution.

The nature of the product obtained will depend in part upon the molar proportion of the various reactants but also upon the pH of the reaction medium. Thus, to prepare the ferric
20 complexes containing a 3:1 molar proportion of ligand: iron(III), the ligand providing compounds and the ferric salt are conveniently mixed in solution in a 3:1 molar proportion and the pH adjusted to a value in the range of 6 to 9, for example 7 or 8. If a similar excess of the compounds:iron is
25 employed but no adjustment is made of the acidic pH which results on the admixture of the compounds and an iron salt such as ferric chloride, then a mixture of 2:1 and 1:1 complexes will instead be obtained. Adjustment of the pH may conveniently be effected by the addition of sodium carbonate as
30 described hereinafter in Example 1. However, a possible alternative, which is of particular interest when preparing the iron complexes in batches of 20 g or more, is to use a hydroxide base such as sodium or ammonium hydroxide.

When using a hydroxide base, the reaction may conveniently be carried out in 4:1 v/v ethanol:water as a solvent and the pH adjusted by the addition of a 2 molar aqueous solution of the base. It will be appreciated that the presence of a proportion
05 of water in the reaction mixture will lead to the retention of a by-product in the iron complex on evaporation of the solvent (a chloride where the iron salt is ferric chloride). However, this can be removed, if desired, by procedures such as crystallisation from a suitable solvent system or sublimation
10 in the particular case of ammonium chloride.

The individual ligand providing compounds may conveniently be used in a 1:1:1 or 2:1 molar proportion, with a 1 molar proportion of the ferric salt, depending on whether all three of the ligands are different or, as is usually preferred, two
15 of the ligands are the same and the third is different. It will be appreciated, however, that the use of such proportions will not lead exclusively to the 1:1:1 or a single 2:1 complex since, although these forms of complex will predominate providing the ligand-providing compounds are of similar
20 reactivity, they will be obtained in admixture with other forms of complex as discussed hereinafter. Indeed, if it is desired to enhance the degree of admixture of different forms of complex which is obtained, the proportions of reactants may be varied to this end. Thus, for example a 1.5:1.5 molar
25 proportion of two different ligands may be used to encourage the formation of a mixture of the two possible types of 2:1 complex differing in the ligand which predominates.

Reaction to form the iron complex is generally rapid and will usually have proceeded substantially to completion
30 after 5 minutes at about 20°C, although a longer reaction time may be used if necessary. Following separation of any precipitated by-product, such as sodium chloride in the case of certain solvent systems, the reaction mixture may conveniently be evaporated on a rotary evaporator to yield the iron complex
35 which will usually be an oil initially which will, however,

often form a glass on standing. The present invention thus further includes a process for the preparation of an iron complex as described hereinbefore, which comprises reacting a mixture of selected hydroxypyridone, hydroxypyrrone and
05 alternative ligand-providing compounds, as described hereinbefore, with ferric ions and isolating the resultant complex.

Whilst for some uses it may be appropriate to prepare the iron complex in a form free from by-products of manufacture
10 apart from other complexes, in other cases, for example with a solid oral formulation as described hereinafter, the presence of by-products such as sodium chloride may be quite acceptable. In general, however, the neutral 3:1 iron(III) complex is of particular interest in a form which is
15 substantially free at least from those by-products which are complexes containing different overall proportions of ligand:iron. Thus the 3:1 complex, although usually obtained and used in a form in which it is in admixture with other types of 3:1 complex as discussed hereinafter, is preferably
20 substantially free from 2:1 and 1:1 complexes. In addition, the isolation of the complex will usually provide it in a form substantially free from any metal-free compound corresponding to a ligand present in the complex. The term "substantially free from" is used herein to indicate the presence of 10% by
25 weight or less of the material referred to.

Certain of the ligand-providing compounds, such as maltol, are available commercially. With others, routes for their preparation are described in the three UK patent applications referred to hereinbefore. Thus, for example, with the
30 hydroxypyrrones a convenient starting material in many instances consists of pyromeconic acid which is readily obtainable by the decarboxylation of meconic acid and may be reacted with an aldehyde to insert a 1-hydroxyalkyl group at the 2-position, which group may then be reduced to produce a 2-alkyl-3-
35 hydroxy-4-pyrone. The preparation of 2-ethyl-3-hydroxy-4-pyrone, etc., by this route is described in US application serial number 310,141 (series of 1960).

It will be appreciated that these are not the only routes available to these compounds and their iron complexes and that various alternatives may be used as will be apparent to those skilled in the art. Moreover, it will be appreciated that
05 certain of the compounds may be converted in vivo to other compounds which are responsible for the metal binding activity observed in vivo. This will be true, for example, of compounds containing ester groups which are likely to be converted to carboxy groups when the compounds are administered orally.

10 The iron complexes of the present invention are of particular interest for several reasons. Firstly, the inclusion in a complex of a mixture of different ligands provides an added dimension to the design of complexes having optimised properties for take up in vivo to provide a
15 controlled supply of iron applicable in a particular human or veterinary context. More specifically, comparative results obtained in human erythrocytes and in non-everted rat jejunal segments suggest that complexes containing mixed ligands may provide a more available source of iron than iron complexes in
20 which the ligands are homogeneous. Secondly, apart from the behaviour of the complexes in vivo, the present invention provides particular advantages in relation to the formulation of iron complexes. In certain contexts, as discussed in more detail hereinafter, liquid formulations of the iron complexes
25 are of particular interest, for example for oral veterinary administration and particularly for parenteral veterinary and human administration, and it has been found that for use in such contexts the solubility of some of the iron complexes of the three previously mentioned UK patent applications is less
30 than might have been desired.

The mixed ligand complexes of the present invention generally show much higher solubilities, both in water and in organic solvents, as compared with those in which the ligands are homogeneous. It is believed that the reason for this lies
35 in the diversity of different stereoisomers of one complex

which can arise when a mixture of ligands is present and which can be augmented by the presence of several different complexes in a reaction mixture obtained from the reaction with iron ions of more than one ligand producing compound. Thus, a 3:1
05 iron(III) complex containing three identical asymmetric ligands can exist in four stereoisomeric forms but when a complex is produced by the reaction of a 3 molar proportion of a mixture of two ligand-producing compounds (A and B) with ferric ions then the following types of complex may be present in the
10 reaction mixture: FeA_3 , FeB_3 , FeAB_2 and FeB_2A . Moreover, although the first two mentioned complexes will exist in four stereoisomeric forms, the last two mentioned complexes will each exist in eight stereoisomeric forms (an even more complex mixture will result if three different ligands are present).
15 It has been found that the four stereoisomers of an FeA_3 or FeB_3 complex will co-crystallise with ease but that for an FeAB_2 or FeB_2A complex the increased number of stereoisomers, and the presence of other 3:1 ligand:iron(III) complexes, prevents such co-crystallisation and ensures that the product
20 is a liquid with enhanced solubility as compared with the usually solid FeA_3 and FeB_3 complexes.

The iron complexes according to the present invention may be formulated for use as pharmaceuticals for both veterinary, for example in an avian or particularly a mammalian context,
25 and human use by a variety of methods. For instance, they may be applied as an aqueous, oily or emulsified composition incorporating a liquid diluent which may often be employed for parenteral administration and therefore may conveniently be sterile and pyrogen free. Oral administration is often
30 preferred for the treatment of iron deficiency anaemia in humans and the complexes of the present invention may be given by such a route. Although compositions incorporating a liquid diluent may be used for oral administration, it is more usual, at least in humans, to use compositions incorporating a solid
35 carrier, for example a conventional solid carrier material such

as starch, lactose, dextrin or magnesium stearate. Such solid compositions may conveniently be of a formed type, for example as tablets, capsules (including spansules), etc.

Although solid compositions may be preferred for the
05 treatment of iron deficiency anaemia in certain contexts, liquid compositions are of interest in other contexts, for example in human and veterinary intramuscular administration and in veterinary oral administration as discussed hereinafter. It is in the area of liquid compositions that the
10 present invention is of particular, although certainly not exclusive, interest. Thus, it is often desirable to produce liquid compositions containing a higher concentration than is readily obtainable with a purely aqueous composition or indeed one containing organic solvents such as simple monohydric
15 alcohols. It has been found with the iron complexes described in the three UK patent applications mentioned hereinbefore that higher concentrations may be achieved by the use of solvents containing two or more hydroxy groups or a hydroxy and an ether group, especially glycols or glycol ethers, either in admixture
20 with water or, for better solubilisation, alone. The glycol ethers of particular interest are the mono-ethers containing as an etherifying group an aliphatic hydrocarbon group of 1 to 6 carbon atoms as described above, for example a methyl group, such a glycol mono-ether being methyl ethylene glycol. In
25 general, however, the glycols themselves are preferred. Examples of such glycols are the simple dihydroxy alkanes such as ethylene glycol as well as those more complex compounds comprising two hydroxy groups attached to a chain containing both carbon and oxygen atoms, such as triethylene glycol, tetraethylene glycol and polyethylene glycol, for example
30 of 4000 daltons molecular weight. Triethylene glycol and especially tetraethylene glycol are of particular interest in view of their very low toxicity. By using such glycols and glycol ethers it is possible to increase solubility for many
35 complexes to 10 to 20 mg/ml. Although such techniques may also

be employed in the formulation of the iron complexes according to the present invention their greater solubility will often allow one with advantage to employ simpler forms of liquid composition and still achieve concentrations considerably in excess of this 10 to 20 mg/ml range.

As indicated, liquid compositions are of particular interest in relation to parenteral administration, a requirement for which arises with humans in certain contexts but also particularly in a veterinary context, for example with pigs. The problems of iron deficiency anaemia in newly born pigs arise primarily during the first three weeks or so of their life when a very rapid weight gain takes place. The usual routes for administration of the iron complexes of the present invention to young piglets are parenteral, for example intramuscular, or oral, for example as a liquid preparation "injected" into the mouth. However, an alternative approach is to enhance the iron content of the milk on which the piglets are feeding by treating the mother pig using oral or parenteral administration, for example with an injectable slow release preparation (such an approach may also be of interest in a human context). When it is applicable to feed piglets on foodstuffs other than the milk of the mother pig, it may also be possible to effect the pharmaceutical administration of the iron complex in this other foodstuff.

Other forms of administration than by injection or through the oral route may also be considered in both human and veterinary contexts, for example the use of suppositories or pessaries for human administration. Another form of pharmaceutical composition of some particular interest is one for buccal or nasal administration and such compositions are discussed hereinafter in more detail.

Compositions may be formulated in unit dosage form, i.e. in the form of discrete portions containing a unit dose, or a multiple or sub-unit of a unit dose. Whilst the dosage of iron complex given will depend on various factors, including the

particular compound which is employed in the composition, it may be stated by way of guidance that maintenance of the amount of iron present in the human body at a satisfactory level will often be achieved using a daily dosage, in terms of the iron
05 content of the compound, which lies in a range from about 0.1 to 10 mg and often in a range of from 0.5 to 10 mg, for example 1 or 2 mg, veterinary doses being on a similar g/Kg body weight ratio. However, it will be appreciated that it may be appropriate under certain circumstances to give daily
10 dosages either below or above these levels. In general, the aim should be to provide the amount of iron required by the patient without administering any undue excess and the properties of the pharmaceutical compositions according to the present invention are particularly suited to the achievement of
15 this aim.

It will be appreciated from the foregoing discussion that more than one iron complex according to the present invention may be contained in the pharmaceutical composition, or indeed, other active compounds may be included in the composition, for
20 example compounds having the ability to facilitate the treatment of anaemia, such as folic acid. Another additional component which may be included in the composition, if desired, is a source of zinc. Iron compounds used in the treatment of iron deficiency anaemia can inhibit the mechanism of zinc
25 uptake in the body and this can cause serious side effects in the foetus when treating anaemia in a pregnant female. It is believed, however, that the iron complexes of the present invention have a further advantage in that they either do not have this effect or exhibit the effect at a lower level than
30 the compounds at present used in the treatment of anaemia. Accordingly, it may often be the case that the level of zinc-providing compound added to the composition may not require to be high or, with preferred formulations of the iron complexes, may be dispensed with altogether.

It has never before been appreciated that the novel iron complexes described herein might be used, and with great advantage, in a pharmaceutical context. Accordingly the present invention includes an iron complex as defined
05 hereinbefore for use in medicine, particularly in the treatment of iron anaemia deficiency. The iron complexes described herein are particularly suited to the treatment of iron anaemia deficiency both in humans and also in a veterinary context, particularly for the treatment of various mammalian species and
10 especially pigs. The complexes will partition into n-octanol indicating that they are able to permeate biological membranes, this property may be confirmed in practice by tests of the ability of the ^{59}Fe labelled iron complexes to permeate erythrocytes. The ability of the compounds in this respect
15 will depend on the nature of the ligands present therein and the reflection of this ability in the K_{part} values of various compounds has been referred to hereinbefore.

Certain aspects of their formulation may enhance the activity of the iron complexes in particular contexts. Thus,
20 although the neutral ferric complexes containing a 3:1 molar proportion of ligand:iron(III) are of particular value as being stable over a wide pH range from about 4 to 5 up to 10, if administered orally they will dissociate at the pH values of less than 4 prevailing in the stomach to form a mixture of
25 the 2:1 and 1:1 complex together with free ligand. Firstly, one of several variations may be employed which avoid or reduce exposure of the iron complex to the acidic conditions of the stomach. Such approaches may involve various types of controlled release system, ranging from one, which may for
30 example be based on a polymer, which simply provides a delayed release of the complex with time, through a system which is resistant to dissociation under acidic conditions, for example by the use of buffering, to a system which is biased towards release under conditions such as prevail in the small
35 intestine, for example a pH sensitive system which is

stabilised towards a pH of 1 to 3 such as prevails in the stomach but not one of 7 to 9 such as prevails in the small intestine. Since the pH of the stomach is higher after a meal, it may be advantageous, whatever method of formulation is used, to administer the iron complexes at such a time.

A particularly convenient approach to a controlled release composition involves encapsulating the iron complex by a material which is resistant to dissociation in the stomach but which is adapted towards dissociation in the small intestine (or possibly, if the dissociation is slow, in the large intestine). Such encapsulation may be achieved with liposomes, phospholipids generally being resistant to dissociation under acidic conditions. The liposomally entrapped 3:1 iron(III) complexes can therefore survive the acid environment of the stomach without dissociating to the 2:1 and 1:1 complexes, and the free hydroxypyrrone. On entry into the small intestine the pancreatic enzymes rapidly destroy the phospholipid-dependent structure of the liposomes thereby releasing the 3:1 complex. Liposome disruption is further facilitated by the presence of bile salts. However, it is usually more convenient to effect the encapsulation, including microencapsulation, by the use of a solid composition of a pH sensitive nature.

The preparation of solid compositions adapted to resist dissociation under acidic conditions but adapted towards dissociation under non-acidic conditions is well known in the art and most often involves the use of enteric coating, whereby tablets, capsules, etc. or the individual particles or granules contained therein, are coated with a suitable material. Such procedures are described, for example, in the article entitled "Production of enteric coated capsules" by Jones in Manufacturing Chemist and Aerosol News, May 1970, and in such standard reference books as "Pharmaceutical Dosage Forms", Volume III by Liebermann and Lackmann (Published by Marcel Decker). One particular method of encapsulation involves the use of gelatine capsules coated with a cellulose acetate

phthalate/diethylphthalate layer. This coating protects the gelatin capsule from the action of water under the acid conditions of the stomach where the coating is protonated and therefore stable. The coating is however destabilised under
05 the neutral/alkaline conditions of the intestine where it is not protonated, thereby allowing water to act on the gelatin. Once released in the intestine the rate of permeation of the intestine wall by the water soluble 3:1 iron(III) complex is relatively constant irrespective of the position within the
10 intestine, i.e. whether in the jejunum, ileum or large intestine. Other examples of methods of formulation which may be used include the use of polymeric hydrogel formulations which do not actually encapsulate the iron complex but which are resistant to dissociation under acidic conditions.

15 A second approach to countering the effect of the acidic conditions prevailing in the stomach is to formulate the iron complexes as a pharmaceutical composition adapted for buccal or nasal administration. Thus, since the buccal and nasal
cavities represent an environment with a pH in the region of 7, a 3:1 neutral iron(III) complex will be taken up by the
20 membranes of the buccal cavity (including the tongue) and the nasal passages in the neutral form, without any significant degree of disproportionation to the corresponding 2:1 and 1:1 complexes. Such a form of administration thus often provides a
25 simple approach to the protection of the 3:1 neutral iron(III) complexes from the strongly acid environment of the stomach.

A further advantage of compositions for buccal or nasal administration is that they provide some form of safety measure as regards overdoses, for example those arising from the taking
30 by children of medication prescribed for adults in the same household. This can pose a considerable problem with existing iron preparations and, although the present iron complexes in any case generally have the advantage of lower toxicity and lower unit dosage levels, the compositions of the present
35 invention provide an added advantage. Thus, it is difficult

rapidly to ingest a large quantity of the iron complex from the buccal cavity or nasal passages and overdosage problems are more likely to arise through swallowing the medicament. If this is done, however, and the iron complex is not formulated
05 in such a way as to protect it from the acid environment of the stomach, disproportionation to the 2:1 and 1:1 complexes will occur thereby significantly reducing the level of iron uptake from the overdose. The generally lower toxicity of the iron complexes will avoid much of the local damage to the
10 gastrointestinal tract which can occur in such circumstances with many commercial iron preparations.

Diluents and carriers suitable for the formulation of compositions for buccal or nasal administration may include various materials in current use in compositions for such
15 administration. Buccal administration is of particular interest and in this case a solid composition is preferred. Such compositions adapted for retention in the mouth rather than swallowing, and consequent release of the active component in the buccal cavity, may take very many forms. These include
20 chewing or bubble gum, lollipops, boiled sweets, effervescent tablets and particularly pastilles and lozenges. Most usually, therefore, the composition will be chewed or sucked to lead to release of the iron complex in the mouth, although it is possible to use tablets, for example in the form of a disc of
25 polymeric material, which are attached to the inside of the buccal cavity and which gradually release the iron complex without being sucked. If desired, liquid compositions may be used in the buccal cavity, particularly aerosol sprays, but these are of less interest. All of these forms of compositions
30 are taken through the mouth but, in contrast to the oral compositions described earlier, are adapted to release of the iron complex in the mouth rather than on being swallowed (although in the process of chewing, sucking etc. a proportion of the iron complex may of course pass into the stomach).

Preferred forms of compositions are pastilles and lozenges and such compositions are sometimes described by the term "linguet", this being a composition suitable for sub-lingual use.

- 05 Specific carriers which may be used in pastilles and lozenges are described in various tests including the British Pharmacopoeia, the British Pharmacopoeia Codex and Martindale, the Extra Pharmacopoeia. One particular example of a base for pastilles is described in the 1980 British Pharmacopoeia and
- 10 consists of a mixture of gelatin, glycerine, sugar, citric acid and amaranth. The rate at which the pastille dissolves in the mouth may be varied as desired with a view to achieving a good level of uptake of iron, the rate of dissolution being reduced, for example, by increasing the proportion of gelatin used.
- 15 Pastilles may conveniently be prepared by forming a melt containing a suitable amount of the iron complex and the carrier and then pouring this into a mould and allowing to dry. One particular example of a base for lozenges is described in the 1959 British Pharmacopoeia Codex and consists
- 20 of a mixture of sucrose, acacia and rose oil water. Once again, the rate of dissolution may be controlled by variation of the ingredients in the base material. Lozenges may conveniently be prepared either by forming a "dough" from which the lozenges are cut or, preferably, by compression. If
- 25 desired, further flavourings can be incorporated in the pastilles or lozenges but the taste of the iron complexes is so acceptable that this may be unnecessary, except perhaps for paediatric formulations.

- Where compositions for nasal administration are employed
- 30 these will usually be liquid and may comprise water and/or suitable organic solvents. Such compositions may conveniently be used either as drops or in the form of an aerosol spray. It is, however, possible to use solid compositions in the form of a snuff if so desired.

A third approach to countering the effect of the acidic conditions prevailing in the stomach, which is described in GB 2117766A and which may also have certain other advantages described therein, is to formulate the iron complex in the
05 pharmaceutical composition together with one or more of the metal-free ligand-providing compounds from which it is derived. Although, this approach may also be considered with the iron complexes of the present invention, particularly when using a metal-free compound of type (1) or (2), it is perhaps
10 rather less attractive in this case in view of the possibility of effecting an exchange between the metal-free compound and the bound ligands, particularly those of type (3), to produce a different type of complex. Also described in GB 2117766A is the use of an iron complex in admixture with a different
15 metal-free iron chelating agent. Once again, this approach may also be considered with the iron complexes of the present invention but is less attractive for similar reasons.

In addition to the pharmaceutical uses of the iron complexes discussed above they are also of potential interest
20 as a source of iron in various other contexts including in cell and bacterial growth, in plant growth, as a colouring agent and in the control of iron transport across membranes.

This invention is illustrated by the following Examples.

EXAMPLES

Example 1

Preparation of iron complexes

- (A) An ethanolic solution of ferric chloride is reacted
05 for 5 minutes at room temperature with a chloroform solution
containing 2 molar equivalents of maltol (3-hydroxy-2-methyl-4-
pyrone) and 1 molar equivalent of 1-ethyl-3-hydroxypyrid-2-
one⁽¹⁾. The resultant solution is neutralised by the addition
of solid sodium carbonate with stirring, the precipitated
10 sodium chloride is removed by filtration and the filtrate is
evaporated to give, as an oil, an essentially quantitative
yield of a mixture of 3:1 complexes in which the (maltol)₂
(1-ethyl-3-hydroxypyrid-2-one) iron(III) complex predominates.
This oil solidifies on standing to give a glass, $\nu_{\text{max}}(\text{glass})$
15 1525, 1560, 1620, 1645 cm^{-1} . (The infra-red spectrum of this
compound, and of the other compounds described hereinafter, is
obtained by dissolving the original glass in chloroform and
evaporating this solution in situ on the sodium chloride plates
to re-form a glass on which the infra-red spectrum is run.)
20 The neutral ferric complexes of (a) ethylpyromeconic acid
(2-ethyl-3-hydroxy-4-pyrone) and maltol and (b) 1-(ethyl-3-
hydroxypyrid-2-one and 1-butyl-3-hydroxypyrid-4-one are
similarly prepared by procedure (A) to give (a) a mixture
of 3:1 complexes in which the (ethylpyromeconic acid)₂ (maltol)
25 iron(III) complex predominates, as a glass, and (b) a mixture
of 3:1 complexes in which the (1-ethyl-3-hydroxypyrid-2-one)₂
(1-butyl-3-hydroxypyrid-4-one) iron(III) complex predominates,

(1) The concentration of the 1 molar equivalent
ligand-providing compound is ca. 0.1M although this
concentration may be varied, for example in a range of 0.1
to 1M, subject only to the solubility of the particular
ligand-providing compounds in the solvent system being used.

as a glass, $\nu_{\max}$ 1525, 1530, 1610 cm^{-1} . (B) An ethanolic solution of ferric chloride is treated with a hot (60°C) ethanolic solution containing 2 molar equivalents of maltol and 1 molar equivalent of ascorbic acid(1) (the use of hot ethanol, rather than cold chloroform as in procedure (A), is required to dissolve the ascorbic acid) and the mixture allowed to cool to room temperature over 30 minutes. The resultant cold solution is neutralised by the addition of solid sodium carbonate with stirring, the precipitated sodium chloride is removed by filtration and the filtrate is evaporated to give, as an oil, an essentially quantitative yield of a mixture of 3:1 complexes in which the (maltol)₂ (ascorbic acid) iron(III) complex predominates. This oil solidifies on standing to give a glass, $\nu_{\max}$ (glass) 1500, 1560, 1600, 1790 cm^{-1} .

The neutral ferric complexes of (a) maltol and leucine and (b) maltol and glycine are similarly prepared by procedure (B) to give (a) a mixture of 3:1 complexes in which the (maltol)₂ (leucine) complex predominates, as a glass, $\nu_{\max}$ 1500, 1560, 1595 cm^{-1} , and (b) a mixture of 3:1 complexes in which the (maltol)₂ (glycine) iron(III) complex predominates, as a glass.

Determination of partition coefficients

The partition coefficient K_{part} , being the ratio (concentration of compound in n-octanol)/(concentration of compound in aqueous phase) on partition between n-octanol and aqueous tris hydrochloride (20 mM, pH 7.4; tris represents 2-amino-2-hydroxymethylpropane 1,3-diol), is measured at 20°C for the 3:1 iron(III) complexes listed in the Table (at 10^{-4}M) by spectrophotometry, the complex prepared as described above being dissolved initially in the aqueous tris hydrochloride. Acid washed glassware is used throughout and, following mixing of 5 ml of the 10^{-4}M aqueous solution with 5 ml of n-octanol for 1 minute, the aqueous n-octanol mixture is centrifuged at 1,000 g for 30 seconds. The two resulting phases are separated for a concentration determination by spectrophotometry on each, the range 340-640 nm being used. Values typical of those obtained are shown in the Table.

TABLE 1
Partition coefficients

Iron(III) complex	Partition coefficient
(maltol) ₂ (1-ethyl-3-hydroxypyrid-2-one)	0.32
(ethylpyromeconic acid) ₂ (maltol)	1.69
(1-ethyl-3-hydroxypyrid-2-one) ₂	3.56
(1-butyl-3-hydroxypyrid-4-one) (maltol) ₂ (ascorbic acid)	0.14
(maltol) ₂ (leucine)	0.17
(maltol) ₂ (glycine)	0.18

Example 2

Donation of iron to apotransferrin

(Maltol)₂ (ascorbic acid) iron(III) was prepared in situ by dissolving maltol and ascorbic acid in a 2:1 molar ratio in an aqueous medium containing morpholine propane sulphonate (MOPS, 25 mM) and sodium hydrogen carbonate (30 mM), the pH being adjusted to 7.4 with hydrochloric acid (50% v/v). To the medium was added a solution of ⁵⁹Fe enriched ferric chloride in 0.1M aqueous hydrochloric acid containing a molar amount of iron equivalent to that of the ascorbic acid. Equal volumes of the solution of the iron complex and of apotransferrin in the same aqueous medium were mixed to provide concentrations in the mixture of 5 x 10⁻⁵M of apotransferrin and 1 x 10⁻⁴M of iron.

The mixture was incubated at room temperature in separate experiments for 10 and 60 minutes and an aliquot was then added to a Sephadex G10⁽¹⁾ column equilibrated with the same aqueous medium described above and the column eluted with that medium,

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fractions being collected and counted. The percentages of ^{59}Fe associated with apotransferrin and with the ligands are calculated as $\text{cpm iron-apotransferrin}/\text{total cpm}$ and $\text{cpm iron-(maltol)}_2 \text{ (ascorbic acid)}/\text{total cpm}$.

05 It was found that after 10 minutes 43% of the ^{59}Fe was associated with apotransferrin and 57% of the ^{59}Fe with the ligands, whilst after 60 minutes the amount of ^{59}Fe donated to the apotransferrin had risen to 75%, only 25% remaining associated with the ligands.

10 Example 3

In vitro tests on permeation of iron complexes into human erythrocytes

 The accumulation of iron by human erythrocytes which are associated with (ethylpyromeconic acid) $_2$ (maltol) iron(III),
15 (maltol) $_2$ (ethylpyromeconic acid) iron(III), (maltol) $_2$ (ascorbic acid) iron(III) and (1-ethyl-3-hydroxypyrid-2-one) $_2$ (1-butyl-3-hydroxypyrid-4-one) iron(III) was studied together with that of a group of four other iron compounds for comparative purposes, this group comprising the 3:1 homogeneous
20 iron complexes, (ethylpyromeconic acid) $_3$ iron(III) and (maltol) $_3$ iron(III), and the salts, ferric NTA (nitrilotriacetic acid) and ferrous sulphate.

 The iron compounds were used in solution in an aqueous medium containing tris (20 mM) and sodium chloride (130 mM),
25 the pH being adjusted to 7.4 with hydrochloric acid (50% v/v). The solid ferric NTA and ferric sulphate, labelled with ^{59}Fe , were dissolved in the aqueous medium to produce a 10^{-4}M solution whilst the iron complexes were prepared in situ by solution of the ligand-providing compound, or of an appropriate
30 molar ratio of ligand-providing compounds, in the aqueous medium and the addition of an appropriate molar proportion of ^{59}Fe enriched ferric chloride in 0.1M aqueous hydrochloric acid, the final concentration of iron being 10^{-4}M .

Packed human erythrocytes (0.5 ml) were incubated in the aqueous medium at 37°C in separate experiments for 5 and 30 minutes in all cases, and for the (maltol)₂ (ascorbic acid) iron(III) complex only, in a series of experiments conducted for 2, 5, 10, 15, 30, 45 and 60 minutes. Following incubation, an aliquot of the erythrocyte incubation medium mixture was separated by centrifugation over two layers of silicone oil ($\rho = 1.07$, and $\rho = 1.2$ respectively. The ⁵⁹Fe levels associated with each of the erythrocytes and the incubation medium were counted and are presented as a distribution ratio (concentration in erythrocytes/concentration in medium) in Table 2 for the 5 and 30 minute experiments and in Table 3 for the other experiment with (maltol)₂ (ascorbic acid) iron(III) (all data in the Tables represents a mean of at least three independent experiments).

It will be seen from Table 2 that ethylpyromeconic acid is more effective as a ligand than maltol in producing uptake of an iron complex by erythrocytes, the uptake both for the complex containing three ethylpyromeconic acid ligands and that containing two of such ligands reaching an equilibrium after only 5 minutes. It should also be noted that a higher level of iron uptake is achieved at 30 minutes for each of the three mixed ligand complexes containing a hydroxypyrrone ligand than for either of the complexes containing three identical hydroxypyrrone ligands.

TABLE 2

Compound	Ligand concentration	Distribution ratio	
		5 minutes	30 minutes
$\text{Fe}^{\text{III}}(\text{ethylpyromeconic acid})_3$	0.3	3.7	3.6
$\text{Fe}^{\text{III}}(\text{maltol})_3$	0.3	0.2	1.1
$\text{Fe}^{\text{III}}(\text{ethylpyromeconic acid})_2$ maltol	0.2:0.1	5.6	5.2
$\text{Fe}^{\text{III}}(\text{maltol})_2$ (ethylpyromeconic acid)	0.2:0.1	1.9	6.3
$\text{Fe}^{\text{III}}(\text{maltol})_2(\text{ascorbic acid})$	0.2:0.1	1.5	5.2
$\text{Fe}^{\text{III}}(1\text{-ethyl-3-hydroxypyrid-2-one})_2(1\text{-butyl-3-hydroxypyrid-4-one})$	0.2:0.1	0.4	1.4
Fe^{III} NTA	0.1	6.2	4.9(1)
$\text{Fe}^{\text{II}}\text{SO}_4$	0.1	5.0	4.3(1)

(1) An analysis of the erythrocyte membranes after incubation shows that these apparently high levels of iron uptake are due to binding to the cell membrane and do not in fact reflect true values of iron uptake.

TABLE 3

[Fe^{III}(maltol)₂-(ascorbic acid)]

Time (minutes)	Distribution ratio
2	6.1
5	8.0
10	10.2
15	12.2
30	17.1
45	26.5 ⁽¹⁾
60	30.8 ⁽¹⁾

(1) These high values arise from dissociation of the complex and donation of iron to intracellular protein.

Example 4

In vitro tests experimentation of non-everted rat jejunal segments by iron complexes

05 The uptake of iron by non-everted rat jejunal segments was studied for (ethylpyromeconic acid)₂ (maltol) iron(III) (maltol)₂ (ethylpyromeconic acid) iron(III) and (1-ethyl-3-hydroxypyrid-2-one)₂ (1-butyl-3-hydroxypyrid-4-one) iron(III). For comparative purposes, similar experiments were carried out with (ethylpyromeconic acid)₃ iron(III) and (maltol)₃ iron(III), the latter complex being used alone and in admixture

10

with an excess of maltol corresponding to either 1 molar proportion or 7 molar proportions of free maltol in addition to the 3 molar proportions contained in the complex in association with 1 molar proportion of iron(III).

05 The iron complexes were prepared in situ in oxygenated HEPES buffer by solution of the appropriate amount of the ligand-providing compound, or of an appropriate molar ratio of ligand-providing compounds, in the buffer and the addition of an appropriate molar proportion of ^{59}Fe enriched ferric
10 chloride in 0.1M aqueous hydrochloric acid, the final concentration of iron being 10^{-4}M .

Rats (male, Sprague Dawley, 70-90 g) were killed and the jejunum removed. It was divided into 3 cm lengths, each being cut lengthwise and opened out. These lengths were further cut
15 into segments of 30-35 mg (3 per incubation flask). The jejunum was incubated at 37°C (pre-incubated flasks being used, each of which contained 3 segments) in a medium of oxygenated HEPES buffer containing the appropriate iron complex for 10 minutes. The tissue and medium were then counted.

20 The uptake of ^{59}Fe was calculated as a distribution ratio (concentration in tissue/concentration in medium) corrected for the water content (80%) and the extracellular fraction (found to be 20% by sulphate space studies) of the tissue. The results are shown in Table 4 and it will be seen, in contrast
25 to the results obtained with erythrocytes, that the uptake of iron from the complex containing three ethylpyromeconic acid ligands is not greater than that for the complex containing three maltol ligands. However as in the erythrocyte experiments, the uptake of iron is greater for the three mixed
30 ligand complexes containing a hydroxypyrrone ligand than for either of the complexes containing three hydroxypyrrone ligands.

TABLE 4

Compound	Ligand concentration (mM)	Distribution ratio
Fe(III)(maltol) ₃ + maltol (1:7 molar ratio of complex:maltol)	1	0.1
Fe(III)(maltol) ₃ + maltol (1:1 molar ratio of complex:maltol)	0.4	0.6
Fe(III)(maltol) ₃	0.3	1.3
Fe(III)(ethylpyromeconic acid) ₃	0.3	1.0
Fe(III)(ethylpyromeconic acid) ₂ (maltol)	0.2:0.1	6.0
Fe(III)(maltol) ₂ (ethylpyromeconic acid)	0.2:0.1	4.3
Fe(III)(maltol) ₂ (ascorbic acid)	0.2:0.1	9.7
Fe(III)(1-ethyl-3-hydroxypyrid-2-one) ₂ -(1-butyl-3-hydroxy-3-methylpyrid-4-one)	0.2:0.1	2.4

CLAIMS

1. An iron complex comprising a ferric cation in combination with at least two different ligands, at least one of which is provided by a compound being:
- 05 (1) 3-hydroxy-4-pyrone or a 3-hydroxy-4-pyrone in which one or more of the hydrogen atoms attached to ring carbon atoms are replaced by an aliphatic hydrocarbon group of 1 to 6 carbon atoms; or
- 10 (2) a 3-hydroxypyrid-2-one or 3-hydroxypyrid-4-one in which the hydrogen atom attached to the ~~nitrogen atom~~ is replaced by an aliphatic acyl group, by an aliphatic hydrocarbon group of 1 to 6 carbon atoms, or by an aliphatic hydrocarbon group substituted by one or more substituents selected from aliphatic acyl, alkoxy, aliphatic amide, aliphatic ester, halogen and hydroxy groups and, optionally, in which one or more of the
- 15 hydrogen atoms attached to ring carbon atoms are replaced by one of said substituents, by an aliphatic hydrocarbon group of 1 to 6 carbon atoms, or by an aliphatic hydrocarbon group substituted by an alkoxy, aliphatic ester, halogen or hydroxy group;
- 20 the trivalent positive charge on the ferric cation being neutralised by the charge on the ligands covalently bound thereto.
2. A neutral 3:1 ligand:iron(III) complex comprising three monobasic, bidentate ligands, of which at least two are
- 25 different, in combination with a ferric cation, each ligand separately being provided by a compound being:
- (1) 3-hydroxy-4-pyrone or a 3-hydroxy-4-pyrone in which one or more of the hydrogen atoms attached to ring carbon atoms are replaced by an aliphatic hydrocarbon group of 1 to 6 carbon
- 30 atoms;
- (2) a 3-hydroxypyrid-2-one or 3-hydroxypyrid-4-one in which the hydrogen atom attached to the nitrogen atom is replaced by an aliphatic acyl group, by an aliphatic hydrocarbon group of 1 to 6 carbon atoms, or by an aliphatic hydrocarbon group

- substituted by one or more substituents selected from aliphatic acyl, alkoxy, aliphatic amide, aliphatic ester, halogen and hydroxy groups and, optionally, in which one or more of the hydrogen atoms attached to ring carbon atoms are replaced by
- 05 one of said substituents, by an aliphatic hydrocarbon group of 1 to 6 carbon atoms, or by an aliphatic hydrocarbon group substituted by an alkoxy, aliphatic ester, halogen or hydroxy group; or
- (3) an alternative compound providing a physiologically
- 10 acceptable, monobasic, bidentate ligand which is capable of bonding covalently to iron; but with the proviso that at least one of the ligands is of type (1) or (2).
3. A complex according to Claim 2, in which the compound of type (3) either contains a first grouping which is an enolic
- 15 hydroxy group or a carboxy group and a second grouping which is an amine group or a hydroxy group, or is a monocarboxylic acid providing an anion containing a grouping $\text{—C}\begin{matrix} \text{O}^- \\ \text{O}^- \end{matrix}$
4. A complex according to Claim 3, in which the compound of type (3) is formic acid, acetic acid, propionic acid, lactic
- 20 acid, gluconic acid, ascorbic acid, glycine, leucine, methionine, phenylalanine, tyrosine, valine or a di- or tri-peptide in which the amino acid residues, which may be the same or different, are residues of amino acids selected from this last mentioned group of six amino acids.
5. A complex according to Claim 4, in which the compound of type (3) is gluconic acid, ascorbic acid, glycine or leucine.
6. A complex according to any of Claims 1 to 5, in which at least two of the ligands are provided by compounds of different types among (1), (2) and (3).
- 30 7. A complex according to any of Claims 1 to 6, in which two of the ligands are provided by the same compound and the other ligand is provided by a different compound.

8. A neutral 3:1 ligand:iron(III) complex comprising three monobasic, bidentate ligands, of which two are the same and the third is different, in combination with a ferric cation, the ligands of both types each separately being provided by a compound being:
- (1) 3-hydroxy-4-pyrone or a 3-hydroxy-4-pyrone in which one or more of the hydrogen atoms attached to ring carbon atoms are replaced by an aliphatic hydrocarbon group of 1 to 6 carbon atoms;
- (2) a 3-hydroxypyrid-2-one or ~~3-hydroxypyrid-4-one~~ in which the hydrogen atom attached to the nitrogen atom is replaced by an aliphatic acyl group, by an aliphatic hydrocarbon group of 1 to 6 carbon atoms, or by an aliphatic hydrocarbon group substituted by one or more substituents selected from aliphatic acyl, alkoxy, aliphatic amide, aliphatic ester, halogen and hydroxy groups and, optionally, in which one or more of the hydrogen atoms attached to ring carbon atoms are replaced by one of said substituents, by an aliphatic hydrocarbon group of 1 to 6 carbon atoms, or by an aliphatic hydrocarbon group substituted by an alkoxy, aliphatic ester, halogen or hydroxy group; or
- (3) formic acid, acetic acid, propionic acid, lactic acid, gluconic acid, ascorbic acid, glycine, leucine, methionine, phenylalanine, tyrosine or valine.
9. A complex according to any of Claims 1 to 8, in which at least one ligand is provided by a compound of type (1).
10. A complex according to Claim 9, in which each ligand separately is provided by a compound of type (1).
11. A complex according to any of the preceding claims, in which the compound of type (1) is 3-hydroxy-4-pyrone or a 3-hydroxy-4-pyrone in which one or more of the hydrogen atoms attached to ring carbon atoms are replaced by an acyclic aliphatic hydrocarbon group of 1 to 4 carbon atoms.

12. A complex according to Claim 11, in which the compound of type (1) is 3-hydroxy-4-pyrone or a 3-hydroxy-4-pyrone in which one or more of the hydrogen atoms attached to ring carbon atoms is replaced by a methyl, ethyl, n-propyl or isopropyl group.
- 05 13. A complex according to Claim 12, in which the substituted 3-hydroxy-4-pyrone has a single substituent at the 2- or 6-position.
14. A complex according to Claim 12, in which the compound of type (1) is selected from the group consisting of 3-hydroxy-4-pyrone, 3-hydroxy-2-methyl-4-pyrone, 3-hydroxy-6-methyl-4-pyrone and 2-ethyl-3-hydroxy-4-pyrone.
- 10 15. A complex according to Claim 12, in which the compound of type (1) is selected from the group consisting of 3-hydroxy-2-methyl-4-pyrone and 2-ethyl-3-hydroxy-4-pyrone.
- 15 16. A complex according to any of Claims 1 to 9, in which at least one ligand is provided by a compound of type (2).
17. A complex according to any of Claims 1 to 8, in which each ligand separately is provided by a compound of type (2).
18. A complex according to any of Claims 1 to 8, 16 and 17, in which the compound of type (2) is a 3-hydroxypyrid-2-one or 3-hydroxypyrid-4-one in which the hydrogen atom attached to the nitrogen atom is replaced by an aliphatic hydrocarbon group of 1 to 6 carbon atoms and, optionally, in which one or more of the hydrogen atoms attached to ring carbon atoms are replaced by the same or a different aliphatic hydrocarbon group of 1 to 6 carbon atoms.
- 20 25 19. A complex according to any of Claims 1 to 8, 16 and 17, in which at least one ligand is provided by a compound of type (1) or a compound of type (2) other than as defined in Claim 18.
- 30 20. A complex according to any of Claims 1 to 8 and 16 to 19, in which the compound of type (2) is a 3-hydroxypyrid-2-one.

21. A complex according to any of Claims 1 to 8, 16 and 17, in which the compound of type (2) is selected from the group consisting of 3-hydroxy-1-methylpyrid-2-one, 1-ethyl-3-hydroxypyrid-2-one, 3-hydroxy-1-propylpyrid-2-one, 3-hydroxy-1-(1'-methylethyl)-pyrid-2-one, 3-hydroxy-1,2-dimethylpyrid-4-one, 1-ethyl-3-hydroxy-2-methylpyrid-4-one, 3-hydroxy-2-methyl-1-propylpyrid-4-one and 3-hydroxy-1-(1'-methylethyl)-2-methylpyrid-4-one.
22. A complex according to Claim 2, being (1-ethyl-3-hydroxypyrid-2-one)₂ (1-butyl-3-hydroxypyrid-4-one) iron(III), (3-hydroxy-2-methyl-4-pyrone)₂ (1-ethyl-3-hydroxypyrid-2-one) iron(III), (3-hydroxy-2-methyl-4-pyrone)₂ (leucine) iron(III), (3-hydroxy-2-methyl-4-pyrone)₂ (glycine) iron(III), (3-hydroxy-2-methyl-4-pyrone)₂ (ascorbic acid) iron(III), (3-hydroxy-2-methyl-4-pyrone)₂ (gluconic acid) iron(III), (3-hydroxy-2-methyl-4-pyrone)₂ (2-ethyl-3-hydroxy-4-pyrone) iron(III) or (2-ethyl-3-hydroxy-4-pyrone)₂ (3-hydroxy-2-methyl-4-pyrone) iron(III).
23. A complex according to Claim 2, being (3-hydroxy-2-methyl-4-pyrone)₂ (2-ethyl-3-hydroxy-4-pyrone) iron(III) or (2-ethyl-3-hydroxy-4-pyrone)₂ (3-hydroxy-2-methyl-4-pyrone) iron(III).
24. A complex according to any of Claims 2 to 23, being substantially free from iron complexes containing one or more of the same ligands in combination with a ferric cation as other than a neutral 3:1 ligand:iron(III) complex.
25. A complex according to any of the preceding claims, being substantially free from a compound in the metal free form corresponding to a ligand present in the complex.
26. A complex according to any of the preceding claims, being substantially free from by-products of manufacture.
27. A pharmaceutical composition comprising an iron complex according to any of Claims 1 to 26, together with a physiologically acceptable diluent or carrier.

28. A pharmaceutical composition comprising one or both of (3-hydroxy-2-methyl-4-pyrone)₂ (2-ethyl-3-hydroxy-4-pyrone) iron(III) and (2-ethyl-3-hydroxy-4-pyrone)₂ (3-hydroxy-2-methyl-4-pyrone) iron(III), together with a physiologically acceptable diluent or carrier.
29. A pharmaceutical composition according to Claim 27 or 28 in solid form.
30. A pharmaceutical composition according to Claim 27, 28 or 29 which is adapted for oral administration.
31. A pharmaceutical composition according to Claim 30 which is formed as tablets or capsules.
32. A pharmaceutical composition according to any of Claims 27 to 31 in delayed release form.
33. A pharmaceutical composition according to any of Claims 27 to 31 which is adapted to release the complex or complexes in the intestine rather than in the stomach.
34. A pharmaceutical composition according to Claim 33, in which the complex or complexes are encapsulated by a material resistant to dissociation under aqueous acidic conditions.
35. A pharmaceutical composition according to Claim 27, 28 or 29 which is adapted for buccal or nasal administration.
36. A pharmaceutical composition according to Claim 35 which is formed as pastilles or lozenges.
37. A pharmaceutical composition according to Claim 27 or 28 in aerosol form.
38. A pharmaceutical composition according to Claim 27 or 28 which contains a sterile, pyrogen-free diluent.
39. A pharmaceutical composition according to any of Claims 27 to 38 in unit dosage form.
40. An animal foodstuff comprising an iron complex according to any Claims 1 to 26.
41. A process for the preparation of an iron complex according to any of Claims 1 to 26 which comprises adding the ligand-providing compounds to ferric ions and effecting reaction therewith.

42. A process according to Claim 41, in which the mixture resulting from the reaction is treated to effect isolation of the iron complexes.
43. An iron complex according to any of Claims 1 to 26 for use
05 in therapy.
44. The use of an iron complex according to any of Claims 1 to 26 for the manufacture of a medicament for use in effecting an increase in the levels of iron in a patient's bloodstream.
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