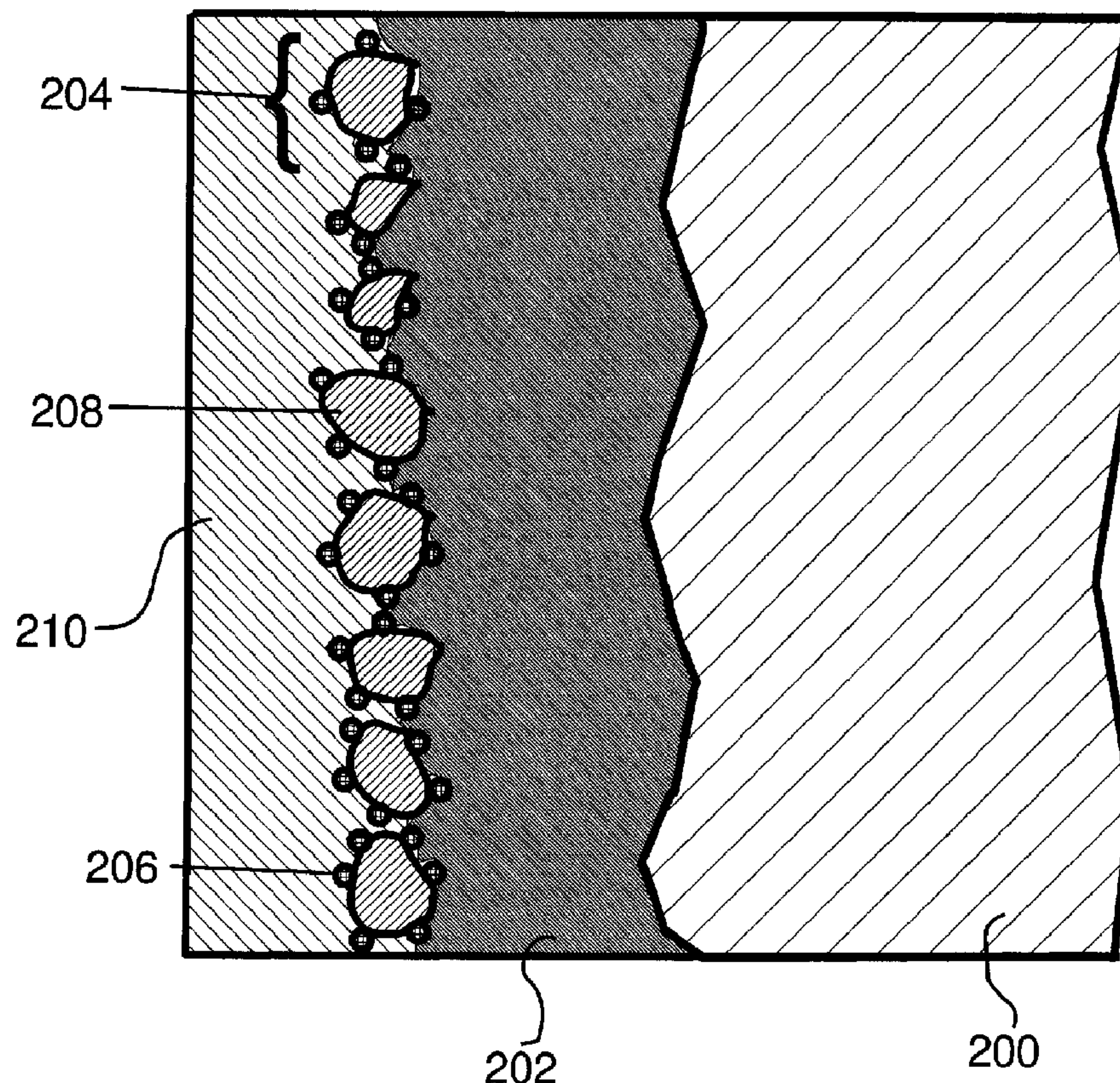




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(54) **Titre :** ELECTROCATALYSEUR, CATHODE DE PILE A COMBUSTIBLE ET PILE A COMBUSTIBLE
(54) **Title:** IMPROVED ELECTROCATALYST, FUEL CELL CATHODE AND FUEL CELL



(57) **Abrégé/Abstract:**

The present invention is related to fuel cells and fuel cell cathodes, especially for fuel cells using hydrogen peroxide, oxygen or air as oxidant. A supported electrocatalyst (204) or unsupported metal black catalyst (206) of cathodes according to an embodiment

(57) Abrégé(suite)/Abstract(continued):

of the present invention is bonded to a current collector (200) by an intrinsically electron conducting adhesive (202). The surface of the electrocatalyst layer is coated by an ion-conducting ionomer layer (210). According to an embodiment of the invention these fuel cells use cathodes that employ ruthenium alloys $\text{RuMe}_I\text{Me}_{II}$ such as ruthenium-palladium-iridium alloys or quaternary ruthenium-rhenium alloys $\text{RuMe}_I\text{Me}_{II}\text{Re}$ such as ruthenium-palladium-iridium-rhenium alloys as electrocatalyst (206) for hydrogen peroxide fuel cells. Other embodiments are described and shown.

ABSTRACT of the Disclosure

The present invention is related to fuel cells and fuel cell cathodes, especially for fuel cells using hydrogen peroxide, oxygen or air as oxidant. A supported electrocatalyst (204) or unsupported metal black catalyst (206) of cathodes according to an embodiment of the present invention is bonded to a current collector (200) by an intrinsically electron conducting adhesive (202). The surface of the electrocatalyst layer is coated by an ion-conducting ionomer layer (210).

According to an embodiment of the invention these fuel cells use cathodes that employ ruthenium alloys $\text{RuMe}_I\text{Me}_{II}$ such as ruthenium-palladium-iridium alloys or quaternary ruthenium-rhenium alloys $\text{RuMe}_I\text{Me}_{II}\text{Re}$ such as ruthenium-palladium-iridium-rhenium alloys as electrocatalyst (206) for hydrogen peroxide fuel cells. Other embodiments are described and shown.

Title: Improved Electrocatalyst, Fuel Cell Cathode and Fuel Cell

by Friedrich W. Wieland

FIELD OF THE INVENTION

[0002] This invention relates to cathodes for fuel cells. More specifically the invention relates to cathodes for fuel cells that use hydrogen peroxide, oxygen or air as oxidants. The invention also relates to semi fuel cells like Magnesium/hydrogen peroxide semi fuel cells that can replace conventional primary batteries.

BACKGROUND

[0003] Fuel cells and semi fuel cells that use hydrogen peroxide or oxygen as oxidant are environmentally friendly methods for generating electricity. They don't produce toxic reaction products during discharge as water is the only product of the use of these oxidants in fuel cells according to the reactions

(1) $\text{H}_2\text{O}_2 + 2 \text{H}_3\text{O}^+ + 2 \text{e}^- \rightarrow 4 \text{H}_2\text{O}$, respectively,

(2) $\text{O}_2 + 4 \text{H}_3\text{O}^+ + 4 \text{e}^- \rightarrow 6 \text{H}_2\text{O}$.

The electrochemical potential calculated from thermodynamic data is +1.77 V for reaction (1) at pH=0, + 1.23 V for reaction (2) but the values reached in practice are significantly lower especially for prior art fuel cell cathodes at current densities of 40-100 mA/cm².

[0004] High performance fuel cells would be a perfect power supply for electrically powered cars as fuel cells can reach an efficiency that is much larger than the efficiency of combustion engines which is limited by the Carnot efficiency $\eta=1-T_1/T_2$ determined by the temperatures T_1 of the cold and T_2 of the hot reservoir and reaches values of typically less than 40% while a fuel cell could reach higher efficiencies. Moreover fuel cells are intrinsically safer than lithium batteries

of the same capacity as only small amounts of educts are present in the fuel cell at the same time and can be stored separately while all highly reactive lithium metal and cathode material, an oxidizer, are mounted next to each other so damage of the separator may result in a violent exothermic reaction of the whole lithium stored in the battery.

[0005] However reaction (1) is rather slow and requires more efficient electrocatalysts in order to reach a low polarization at current densities of 10 mA/cm^2 and above. Cathodes with prior art electrocatalysts still cannot reach low polarizations at current densities of 100 mA/cm^2 and above and suffer from other disadvantages like strong hydrogen peroxide decomposition.

[0006] In spite of substantial efforts to develop improved oxygen and hydrogen peroxide cathodes for fuel cells during the last five decades the power density that could be reached by such fuel cells is still fairly limited as the polarization of the cathodes is already quite large at rather small current densities due to small value of the exchange current densities j_0 for the above reactions (1) and especially (2).

[0007] In addition rather large amounts of very expensive catalysts like platinum and platinum alloys have to be used in order to reach current densities required for an electrically powered car as the catalyst utilization is quite low (about 9% for typical PEM-fuel cells). An estimate for the manufacturing costs of the electrodes of a fuel cell for an electrically powered car was \$50-100 per kW according to S. Srinivasan ("Fuel Cells", Springer, 2006, p. 603). For an electrically powered car with the performance of conventional cars (80 kW power) manufacturing costs of \$4000-\$8000 for the electrodes alone would be therefore expected.

[0008] State of the art fuel cell electrodes for polymer electrolyte membrane (PEM) fuel cells are produced by a coating process using an ink of catalyst mixed with a dispersion of fluoropolymer ionomer like copolymers of tetrafluoroethylene and perfluorovinylether sulfonic acid commonly sold under the trademark "NAFION" by E.I. DuPont de Nemours and Company, Wilmington, Delaware. Such an electrode is shown in Fig. 1. The supported catalyst (104) with platinum, palladium or iridium electrocatalyst centers (106) is randomly distributed in the catalyst-ionomer layer (102) formed from the ink on a conducting current collector (100). Up to now in spite of tremendous research efforts over many decades researchers didn't recognize the disadvantages that arise from this random electrode structure.

[0009] "NAFION" is an ion conductor that is not electron-conducting. But in order to act as an electrocatalyst a catalyst particle must take up electrons from the current collector of the cathode as it is shown by the arrows in Fig. 1 illustrating the flow of electrons within the cathode.

Therefore catalyst utilization is reduced by the random dispersion of the catalyst in a non-electron-conducting polymer as only the fraction of the catalyst that is in electrical contact to the current

collector is acting as electrocatalyst for production of electrical energy.

[0010] Moreover prior art catalysts like platinum, palladium-iridium or gold for hydrogen peroxide cathodes according to reaction (1) show strong polarization at rather small current densities of 10 mA/cm². According to the literature magnesium/hydrogen peroxide-semi fuel cells (open circuit voltage 2.1 V) with conventional cathodes can deliver only a voltage of 1.3 V at current densities of 40 mA/cm² and 25 ml/min flow rate. The situation is similar for oxygen cathodes according to reaction (2) due to the very low exchange current density j_0 .

[0011] Besides efficient prior art electrocatalysts like palladium-iridium (50 atomic-%) cannot be used in concentrated catholyte solutions comprising hydrogen peroxide ($c(\text{H}_2\text{O}_2) > 0.5$ mole/l) that would be useful for high power density fuel cells that operate at high current densities because of decreasing efficiency of prior art electrocatalyst palladium-iridium (50 at.-% Ir) at $c(\text{H}_2\text{O}_2) > 0.25$ mole/l for reaction (1). This prior art electrocatalyst generates much oxygen by catalytical hydrogen peroxide decomposition according to (3) $2 \text{H}_2\text{O}_2 \rightarrow 2 \text{H}_2\text{O} + \text{O}_2$. The energy density decreases from over 700 Wh/kg (for $c(\text{H}_2\text{O}_2) = 0.03$ mole/l) to about 400 Wh/kg (for $c(\text{H}_2\text{O}_2) = 0.25$ mole/l) because of this parasitic reaction instead of an expected increase due to the reduced mass of the catholyte because of the reduced water content in the catholyte as a result of the increased hydrogen peroxide concentration.

[0012] Information relevant to attempts to address these problems can be found in U.S. Patent Applications No. 2008/0182153 A1, 2008/0193827 A1, 2008/0063922 A1, 2008/0054226 A1, 2004/0224218 A1, 2004/0191605, U.S. Patent Nos. 7175930, 5296429, 5445905, 6465124 and the articles Electrochemistry Communications 10 (2008), 1610, in print, Journal of Power Sources 165 (2007), 509 and Journal of Power Sources 164 (2007), 441.

[0013] However, each one of these references suffers from one or more of the following disadvantages as long diffusion paths for educts (the oxidants H_2O_2 or O_2 and H_3O^+) and products (H_2O) from the electrolyte to the electrocatalyst and vice versa, limited durability of electrodes, high costs of the catalysts, high manufacturing costs due to complicated manufacturing processes, strong decomposition of hydrogen peroxide at the surface of the catalyst and low utilization efficiency of hydrogen peroxide, impracticality of the use of concentrated solutions of hydrogen peroxide, strong polarization at large current densities and low utilization of the catalyst due to a missing conduction path for electrons.

[0014] For the foregoing reasons, there is a need for hydrogen peroxide cathodes and oxygen cathodes for fuel cells that are more efficient, less expensive to manufacture and durable and that can deliver higher current densities with lower polarizations and that can be operated in

concentrated solutions of hydrogen peroxide.

SUMMARY

[0015] The present invention is directed to fuel cell cathodes that satisfy this need.

Fuel cell cathodes having features of the invention comprise an electrode structure that has an optimized topology for the transport of educts and products of the cathode reaction (1) or (2) as well as the transport of electrons involved in the cathode reaction.

[0016] A fuel cell cathode according to an embodiment of the invention comprises an electrocatalyst that is bonded to the current collector using an adhesive that is intrinsically conducting electrons. This increases the number of catalyst particles of the electrocatalyst layer that are electrically connected to the current collector by the electron-conducting adhesive. Therefore the catalyst utilization is larger than in prior-art electrodes using random distribution of catalyst particles in a polymer that does not conduct electrons.

[0017] The catalyst layer is coated by an ion conducting layer of an ionomer such as a copolymer of tetrafluoroethylene and a perfluorovinyl ether sulfonic acid commonly sold under the trademark "NAFION" by E.I. DuPont de Nemour and Company, Wilmington, Delaware. This thin ionomer layer ensures efficient transport of educts of the cathode reaction like the oxidants H_2O_2 or O_2 and H_3O^+ and product (water) by diffusion. The ionomer layer also protects the catalyst layer against abrasion and improves durability of the cathode and wettability of the catalyst layer. This topology ensures optimum electronic and ionic conductivity as well as optimum transport of oxidizer and water.

[0018] The catalyst can be a supported electrocatalyst (such as platinum on carbon) or an unsupported catalyst (such as platinum black) and consist of any platinum metal, transition metal, bismuth, tin or aluminum or alloys thereof.

[0019] For hydrogen peroxide cathodes ruthenium or ruthenium based alloys are preferred according to an embodiment of the invention as ruthenium and ruthenium-based alloys are more effective electrocatalysts for the electrochemical reduction of hydrogen peroxide in concentrated solutions ($c > 1 \text{ mol/l}$) than prior art electrocatalysts. Ruthenium catalyzed hydrogen peroxide cathodes have a more positive open cell potential and can therefore deliver a higher open cell voltage in a fuel cell than hydrogen peroxide cathodes using prior art electrocatalysts. Besides the polarization of hydrogen peroxide cathodes using ruthenium alloy electrocatalysts in 2.32 M H_2O_2 solutions is lower than the polarization of hydrogen peroxide cathodes using other electrocatalysts while the rate of generation of oxygen by catalytical hydrogen peroxide decomposition (3) $2 \text{ H}_2\text{O}_2 \rightarrow 2 \text{ H}_2\text{O} + \text{O}_2$ is significantly lower for ruthenium-based electrocatalysts than for other prior art

electrocatalysts. Moreover durability of ruthenium-based electrocatalysts in hydrogen peroxide is excellent. In addition they are considerably less expensive than other platinum metals that are used as prior-art electrocatalysts.

[0020] According to an embodiment of the present invention binary ruthenium alloys RuMe_I with small amounts (2 at.-%) of metal Me_I selected from the group consisting of palladium, iridium, rhenium, platinum, osmium, and rhodium are preferred as electrocatalysts for hydrogen peroxide cathodes. Alloys with Me_I selected from the group consisting of palladium, iridium, and rhenium are more preferred. Ternary alloys $\text{RuMe}_I\text{Me}_{II}$ with Me_{II} selected from the group consisting of palladium, iridium, rhenium, platinum, osmium, and rhodium with Me_{II} different from Me_I such as ruthenium-palladium-iridium, ruthenium-palladium-rhenium, and ruthenium-iridium-rhenium using a small amount of rhenium as additive are most preferred, quaternary alloys ruthenium-palladium-iridium-rhenium are optimum.

[0021] According to a further embodiment of the present invention electrocatalysts selected from the group consisting of platinum, ruthenium, rhodium, osmium, rhenium, palladium, iridium, chromium, cobalt, nickel, manganese, vanadium, silver, titanium, tungsten, aluminum, tin, silicon, molybdenum, bismuth, and alloys thereof are preferred electrocatalysts for oxygen cathodes.

[0022] According to a version of the invention the material of the current collector is chosen that is resistant against corrosion by the electrolyte and can consist of carbon paper, carbon fiber fabric, titanium, or conducting polymers.

[0023] The preferred electron conducting adhesive comprises an intrinsically electron conducting polymer like polyaniline and a dopant. According to an embodiment of the invention an ionomer such as "NAFION" may be used as a dopant in an acidic catholyte to obtain an adhesive of excellent electrical conductivity if the electrode is wetted by a electrolyte such as dilute sulfuric acid. Besides this adhesive offers good ionic conductivity for hydronium ions and allows transport of the oxidant hydrogen peroxide by diffusion.

[0024] According to an embodiment of the invention the ruthenium electrocatalyst or ruthenium-based alloy electrocatalyst can be deposited after bonding the supporting carbon by a process comprising a step of coating the cleaned electrode by electroplating or by electroless plating on a electrode that is cleaned and catalyzed by deposition of palladium atoms. A plating bath that comprises a ruthenium nitridochloro complex is preferred for electroplating.

[0025] For oxygen cathodes various cathode designs are presented that maximize surface area and oxygen diffusion.

[0026] According to an embodiment of the invention the electron conducting adhesive composition comprises an intrinsically electron-conducting polymer like polyaniline, a

fluoropolymer ionomer like "NAFION" and a solvent like dimethyl sulfoxide, dimethyl formamide or N-methyl pyrrolidinone. According to an embodiment of the invention an intrinsically electron conducting pressure sensitive adhesive (PSA) further comprising an elastomer and a tackifier is preferred as electron conducting adhesive.

BRIEF DESCRIPTION OF THE DRAWINGS

[0027] These and the other features, aspects and advantages of the present invention will become better understood with regard to the following description, appended claims and accompanying drawings, where:

[0028] Fig. 1 shows a sectional view of a portion of a prior-art fuel cell electrode with flow of electrons from the current collector to the catalyst particles active as electrocatalyst;

[0029] Fig. 2A shows a sectional view of a portion of a hydrogen peroxide cathode as an embodiment of the present invention;

[0030] Fig. 2B shows a sectional view of a particle of a supported electrocatalyst with ruthenium alloy metal centers;

[0031] Fig. 2C shows a sectional view of a portion of a hydrogen peroxide cathode with unsupported electrocatalyst as a further embodiment of the present invention;

[0032] Fig. 2D shows a sectional view of an unsupported metal black electrocatalyst particle;

[0033] Fig. 3A shows a sectional view of an oxygen cathode as a further embodiment of the present invention;

[0034] Fig. 3B shows a sectional view of a coated carbon fiber of the oxygen cathode shown in Fig. 3A;

[0035] Fig. 3C shows a sectional view of a coated bundle of carbon fibers according to an alternative embodiment of the oxygen cathode;

[0036] Fig. 3D shows an oxygen cathode of Fig. 3A comprising coated carbon fibers shown in Fig. 3C attached to a PEM membrane of a polymer electrolyte fuel cell;

[0037] Fig. 3E shows a larger portion of a carbon fiber coated cathode of Fig. 3A;

[0038] Fig. 4A shows a portion of an electrocatalyst-coated porous conducting body of an oxygen cathode according to a further embodiment of the present invention;

[0039] Fig. 4B shows two pores of an electrocatalyst-coated porous conducting body of an oxygen cathode of Fig. 4A;

[0040] Fig. 4C shows a side view of the pores of Fig. 4B;

[0041] Fig. 4D shows a section through the pores of Fig. 4C;

[0042] Fig. 5 shows a fuel cell according to an embodiment of the present invention;

[0043] Fig. 6 shows a polarization curve for smooth bright films of the ruthenium alloy $\text{Ru}_{0.96}\text{Pd}_{0.02}\text{Ir}_{0.02}$ (96 at.-% ruthenium, 2 at.-% palladium, 2 at.-% iridium) on nickel sheets compared to bright films of a palladium-iridium alloy (50 at.-% Ir) on a nickel sheet as electrocatalyst for hydrogen peroxide cathodes according to a further embodiment of the present invention in a static solution of 3.2 M H_2O_2 , 0.4 M H_2SO_4 at 20°C;

[0044] Fig. 7 shows a comparison of prior art electrocatalysts platinum and iridium to ruthenium in 0.01M H_2O_2 /0.5M H_2SO_4 . Ruthenium appears to be a less active electrocatalyst under these conditions at current densities $>0.4 \text{ mA/cm}^2$;

[0045] Fig. 8 shows a polarization curve of a cathode according to the present invention comprising ruthenium-palladium-iridium (2 at.-% Pd, 2 at.-% Ir) on etched "VULCAN XC72R" carbon black bonded to Toray TGP-H-120 carbon paper using an intrinsically conducting adhesive. This measurement was performed in a static electrolyte of 2.3 M hydrogen peroxide/0.5M sulfuric acid without stirring against a magnesium anode at 20°C;

[0046] Fig. 9 shows a polarization curve of a ruthenium cathode in an electrolyte of H_2O_2 /1 M perchloric acid with and without an 1.2 M sulfuric acid additive in order to demonstrate the influence of formation of Caro's acid H_2SO_5 on the polarization curve;

[0047] Fig. 10 shows a comparison of the most preferred ternary ruthenium alloy Ru 96 at.-% Pd 2 at.-% Ir 2 at.-% to binary ruthenium alloy Ru 98 at.-% Pd 2 at.-%;

[0048] Fig. 11 shows a sectional view of a bipolar electrode comprising a cathode according to the present invention and an anode; and

[0049] Fig. 12 shows a sectional view of a membrane electrode assembly comprising a cathode according to the present invention, a polymerelectrolyte membrane and an anode.

[0050] These and the other features, aspects and advantages of the present invention are better understood with respect to the following description and appended claims.

DETAILED DESCRIPTION

[0051] As shown in Fig. 2A a cathode according to an embodiment of the present invention comprises a current collector (200) that is coated by a layer of an adhesive (202) that is conducting electrons. The adhesive (202) bonds particles of the electrocatalyst (204, cf. Fig. 2B) to the current collector. This ensures good electrical contact of the electrocatalyst to the current collector that is required in order to increase the utilization of the electrocatalyst. The adhesive may be applied by screen printing, painting or other techniques used to apply an adhesive.

[0052] As electrocatalyst a supported catalyst like platinum metals (206) on carbon (208)

or an unsupported catalyst like platinum black may be used.

[0053] The cathode further comprises an ion-conducting coating (210) of the catalyst particles that consists of an ionomer. This coating ensures good ionic conductivity at the surface of the catalyst particles and good wetting of the cathode. It also ensures efficient transport of educts of the cathode reaction like H_2O_2 or O_2 and H_3O^+ -ions to the electrocatalyst and transport of the reaction product water from the electrocatalyst to the electrolyte.

[0054] The ionomer coating also increases durability of the cathodes as it also acts like a polymer coating reinforcing catalyst particles in the catalyst layer.

[0055] The topology of a cathode according to an embodiment of the invention as shown in Fig. 2A thus optimizes the above conductivity and minimizes diffusion problems that occur in prior art fuel cell cathodes. Besides this topology also increases utilization of the catalyst as more catalyst particles are electrically connected to the current collector which helps to reduce catalyst loading and costs. Moreover it allows use of high surface area supported electrocatalysts and metal black electrocatalysts with a high surface area that minimizes polarization.

[0056] According to an embodiment of the present invention the current collector (200) can consist of carbon paper, carbon fiber fabric, titanium mesh or titanium meshed metal baffle and pressed carbon rovings with a binder but other material might be suitable, too. The current collector (200) can also comprise conducting polymers such as polyanilines, polythiophenes, and polypyrroles.

[0057] Carbon paper like the material commonly sold under the trademark "Toray TGP-H-060" (190 μm thickness) or "TGP-H-120" (370 μm thickness) by Toray Industries America Inc., New York, NY is preferred. The latter TGP-H-120 is most preferred for fuel cells that are subject to stronger mechanical stress like in automotive applications. Carbon fiber fabric, pressed carbon fiber rovings and titanium mesh or titanium meshed metal baffle are preferred less expensive alternatives of good conductivity and excellent corrosion resistance in electrolytes containing acids and hydrogen peroxide.

[0058] As electron-conducting adhesive (202) an adhesive is preferred that comprises an intrinsically electron-conducting polymer. This is advantageous compared to conventional conducting adhesives that contain graphite or silver particles that have a size similar to or larger than the size of the electrocatalyst particles (204). A conducting adhesive that is intrinsically conducting electrons can provide electrical contact even to very small catalyst particles that are much smaller than the conducting graphite or silver particle additives of conventional conducting adhesives embedded in a non-conducting matrix. Therefore the use of an intrinsically conducting adhesive ensures electrical contact of an increased number of electrocatalyst particles (204) to the current collector (200) as even a very small adhesive link can provide an electrical connection to the

current collector (see Fig. 2) that would consist of non-conducting polymer in a conventional conducting graphite- or silver-based adhesives. Besides conventional adhesives that contain silver would dissolve in an electrolyte of a fuel cell that uses an acidic hydrogen peroxide solution as oxidizer.

[0059] As intrinsically conducting polymer of the conducting adhesive (202) a large number of polymers are suitable. Polyaniline (PANI) is a preferred intrinsically conducting polymer as it is less expensive and commonly available from a number of manufacturers like Ormecon Chemie GmbH & Co. KG, D-22941 Ammersbeck, Germany. Besides polyaniline is stable in electrolytes that contain strong oxidizers like hydrogen peroxide and it offers excellent conductivity.

Nevertheless a large number of copolymers of aniline with derivatives of aniline such as anthranilic acid or monomers of other conducting polymers such as thiophene, pyrrole, furane, as well as a large number of other intrinsically conducting polymers like polythiophenes, for example poly(3,4-ethylenedioxythiophene) (PEDOT), polypyrroles, polyfuranes, polyparaphenylenes, polyazulenes, polyindoles, polypyridines, polypyrazines, polytriazines, polythiazoles, polyimidazoles, polyquinolines, polybenzimidazoles, polytriazoles, polyoxydiazoles, polythianaphthenes, polycarbazoles, polybenzothiophenes, polybenzofuranes, polyheptadiyne, polyparaphenylene, and polyparaphenylene vinylene and their substituted derivatives and copolymers may be also suitable intrinsically conducting polymers for the conducting adhesive (202).

[0060] Besides copolymers of conducting polymers like polythiophenes with acrylates can be used as intrinsically electrically conducting polymers in the adhesive (202). For example the adhesive can comprise doped poly(3,4-ethylenedioxythiophene) (PEDOT) that is tetramethacrylate end-capped which is commonly sold under the trademark "OLIGOTRON" by TDA Research Inc., Wheat Ridge, Colorado, as a solution in nitromethane. Such adhesives can be cured thermally or by ultraviolet light. Thermal curing is preferred.

[0061] The adhesive (202) may also comprise a dopant if the polymer is not self-doped like polyaniline-polyanthranilic acid copolymers. Alternatively doped electron-conducting polymers (for example PANI emeraldine salts) can be used. As dopants Brønsted acids like hydrochloric acid, sulfuric acid, perchloric acid as well as carboxylic acids or sulfonic acids like methanesulfonic acid, p-toluenesulfonic acid, dodecylbenzene sulfonic acid (DBSA), dinonyl naphthalene sulfonic acid, camphor sulfonic acid (CSA), or polymeric acids like poly(styrenesulfonic acid) as in PEDOT:PSS commonly sold under the trademark "CLEVIOS P" by H.C. Stark, GmbH, D-38642 Goslar, Germany, can be used. Alternatively a polymer like polyaniline can be N-alkylated in order to produce conducting N-alkyl-polyanilines.

[0062] Instead of conventional dopants ionomers that have acidic groups like sulfonic acid groups can be used as dopants. This is more preferred as ionomers are also ion-conducting.

An adhesive (202) that comprises an ionomer is therefore electron-conducting and ion-conducting at the same time. Copolymers of tetrafluoroethylene and a perfluorovinyl ether sulfonic acid that are commonly sold under the trademark “NAFION” by E.I. DuPont de Nemours and Company are most preferred ionomers as dopants for the intrinsically conducting adhesive (202). An adhesive that comprises a solution of polyaniline (emeraldine base) and “NAFION” in an organic solvent provides a suitable bonding strength and excellent electron-conductivity when immersed in the acidic electrolyte of the fuel cell. The adhesive may be further comprising a solvent for polyaniline (emeraldine base) such as dimethyl sulfoxide (DMSO), dimethyl formamide and N-methyl pyrrolidinone although other solvents might be suitable, too. More preferred are dispersions of PANI and “NAFION” in mixtures of DMSO and alcohols (such as propanol) and water. Preferred are also solutions of PEDOT:PSS and “NAFION” in alcohol/water as intrinsically electron conducting adhesives (202).

[0063] Intrinsically conducting pressure sensitive adhesives are most preferred conducting adhesives as they simplify production of cathodes. A polyaniline-dopant complex like PANI-dodecylbenzene sulfonic acid (PANI-DBSA) that is soluble in an organic solvent like p-xylene can be used for production of an intrinsically conducting PSA further comprising an elastomer and a tackifier like poly- α -pinene.

[0064] As electrocatalyst (206) any transition metal catalyst, any main group element and all alloys thereof may be used. For oxygen or air cathodes platinum, silver, nickel are well known effective prior art catalysts. Recently alloys of platinum with chromium, vanadium, titanium, tungsten, aluminum, tin, silicon, nickel, cobalt, iron, manganese, and molybdenum were developed for cathodes of hydrogen-oxygen fuel cells.

[0065] For cathodes using hydrogen peroxide as oxidant platinum, palladium, iridium, gold, silver, niobium, nickel, nickel-aluminum, titanium, titanium boride, iridium oxide, glassy carbon, porphyrine complexes, peroxidase, cobalt, tungsten, bismuth and palladium-iridium alloys were extensively tested in prior art references as electrocatalysts. Palladium-iridium nanoparticles and binary as well as ternary palladium, iridium, cobalt, tungsten, bismuth, and molybdenum alloys were tested recently as more or less effective catalysts.

[0066] According to an embodiment of this invention ruthenium, osmium, and rhenium are effective catalysts for hydrogen peroxide cathodes in concentrated hydrogen peroxide catholytes. The latter two metals (Os, Re) dissolve in acidic hydrogen peroxide solutions as the electrode potentials Re/Re^{3+} ($E^0=+0.30\text{V}$) and Os/OsO_4 ($E^0=+0.838\text{V}$) are lower than the potential of H_2O_2

($E^0=+1.77\text{V}$). Thin films of osmium and rhenium dissolve within a few seconds. It can be expected that technetium would be also a more or less effective catalyst that might dissolve under these conditions, too. Therefore these metals rhenium and osmium may be used as alloys with other noble metals that are corrosion resistant in such catholytes.

[0067] It is surprising that ruthenium does not dissolve in acidic concentrated hydrogen peroxide solutions in spite of an electrode potential of Ru/Ru^{2+} of only $E^0=0.455\text{V}$ as Ruthenium anodes readily dissolve in dilute (0.5 M) sulfuric acid during anodic polarization under formation of Ruthenium tetroxide (RuO_4). It was discovered that ruthenium metal is stable and doesn't dissolve in 2.32 M solutions of hydrogen peroxide in sulfuric acid and that ruthenium is a more efficient electrocatalyst for cathodes of fuel cells using concentrated hydrogen peroxide as oxidant than prior art electrocatalysts palladium, iridium and significantly superior to palladium-iridium alloys regarding catalyst stability and parasitic hydrogen peroxide decomposition. This excellent electrocatalytical activity in 2.32 M H_2O_2 solution is quite surprising as it was found that ruthenium is an inferior electrocatalyst in a dilute 0.01 M H_2O_2 /0.5 M H_2SO_4 catholyte (that was typically used in the literature for evaluation of electrocatalysts for hydrogen peroxide reduction) compared to prior art electrocatalysts such as the more efficient platinum (see Fig. 7) or iridium or the less efficient palladium and the most efficient prior-art electrocatalyst palladium-iridium. Alloys of palladium with ruthenium showed only little advantage compared to palladium in a catholyte comprising 0.03 M H_2O_2 .

[0068] As most prior art electrocatalysts for hydrogen peroxide reduction are also strongly catalyzing hydrogen peroxide decomposition according to reaction (3) it is surprising that ruthenium and alloys of ruthenium show a significantly reduced parasitic decomposition of hydrogen peroxide and oxygen evolution in spite of excellent electrocatalytical activity.

[0069] It was found that ruthenium alloys RuMe_1 with Me_1 selected from the group consisting of palladium, iridium, rhenium, platinum, osmium, and rhodium are superior to pure ruthenium metal as electrocatalyst for fuel cells using concentrated hydrogen peroxide solutions with $c(\text{H}_2\text{O}_2)>1\text{ mol/l}$. Preferred metals Me_1 are palladium, iridium, and rhenium.

[0070] It is surprising that even small amounts of alloy component Me_1 of palladium, iridium or rhenium in the range of 10 ppm start to reduce polarization at current densities of $j>10\text{ mA/cm}^2$. Contents of metals Me_1 of less than 50% are preferred because of the higher open cell voltage, contents of less than 5% are most preferred. In spite of the small content of Me_1 such alloys have a significantly higher electrocatalytical activity than ruthenium. Alloys RuMe_1 with 1-2 at.-% Me_1 are optimum. Moreover rhenium as additive Me_1 increases the open cell voltage of fuel cells using such a hydrogen peroxide cathode.

[0071] It was found that particular ternary ruthenium alloys $\text{RuMe}_I\text{Me}_{II}$ with $\text{Me}_I, \text{Me}_{II} \in \{\text{Pd}, \text{Ir}, \text{Re}, \text{Pt}, \text{Os}, \text{Rh}\}$, $\text{Me}_I \neq \text{Me}_{II}$ are even particularly superior to the above binary alloys RuMe_I electrocatalysts especially at current densities $j > 20 \text{ mA/cm}^2$ (see Fig. 10). Ternary ruthenium alloys $\text{RuMe}_I\text{Me}_{II}$ with $\text{Me}_I = \text{palladium}$, $\text{Me}_{II} = \text{iridium}$ (i.e. RuPdIr) are most preferred, quaternary ruthenium alloys further comprising rhenium $\text{RuMe}_I\text{Me}_{II}\text{Re}$ are optimum electrocatalysts for hydrogen peroxide fuel cell cathodes.

[0072] Fig. 6 shows a comparison of a polarization curve for a nickel sheet coated by a bright smooth layer of a ruthenium-palladium-iridium alloy (96 at.-% Ru, 2 at.-% Pd, 2 at.-% Ir) in an electrolyte of 2.32 M H_2O_2 in 0.4 M H_2SO_4 as cathode compared to a nickel sheet coated with a bright palladium-iridium alloy layer (50 at.-% iridium) that was preferred in most prior art publications about hydrogen peroxide fuel cell cathodes. It is evident that the ruthenium-palladium-iridium-alloy is a more efficient electrocatalyst for hydrogen peroxide cathodes. Ruthenium and ruthenium-based alloys are therefore preferred catalysts for fuel cells using hydrogen peroxide as oxidant as they are also generating much less oxygen by catalytic hydrogen peroxide decomposition (3) compared to palladium-iridium (50% at.-% Pd) and have a better durability than thin palladium-iridium films that significantly lose electrocatalytical activity after a few minutes of use in 2.32 M H_2O_2 (see table 1).

[0073] Table 1 shows properties of ruthenium and palladium-iridium (50-at.-% Ir) films on a nickel sheet in a solution that consisted of 2.32 M H_2O_2 for measurement of oxygen generated as by-product by catalytical hydrogen peroxide decomposition at 21°C (measured volume converted to $T=273 \text{ K}$, $p=1013.25 \text{ hPa}$ by calculation) and results of a test of durability of electrocatalytic activity in a catholyte that consisted of 2.32 M H_2O_2 and 0.4 M H_2SO_4 . Ruthenium and ruthenium alloys generate only about 1/23 (respectively 1/19) of the amount of oxygen generated by hydrogen peroxide decomposition on palladium-iridium alloy films. Therefore ruthenium coated cathodes are preferred in catholytes of $c(\text{H}_2\text{O}_2) > 2 \text{ mol/l}$.

Table 1		
electrocatalyst	Amount of generated oxygen by catalytical hydrogen peroxide decomposition [ml/(cm ² s)]	Durability of electrocatalytical activity in 2.32 M H ₂ O ₂ and 0.4 M H ₂ SO ₄
ruthenium	small (0.0068)	good
ruthenium-palladium (2 at.-% Pd, Ru balance)	small (0.0082)	good
ruthenium-palladium-iridium (2 at.-% Pd, 2 at.-% Ir, Ru balance)	small	good
palladium-iridium (50at.-% Ir)	large (0.155)	poor

[0074] The use of ruthenium as electrocatalyst for fuel cell cathodes using hydrogen peroxide as oxidant also reduces costs as the ruthenium price is considerably lower than the price of iridium and palladium which are preferred for prior art cathodes.

[0075] Ruthenium-based alloys are preferred as electrocatalysts for fuel cells employing concentrated hydrogen peroxide ($c(\text{H}_2\text{O}_2) > 1 \text{ mol/l}$) as oxidant. Especially alloys RuMe_I with palladium, iridium, platinum, and osmium have shown a decrease of polarization for current densities of $50\text{-}100 \text{ mA/cm}^2$ in concentrated solutions and are more preferred, alloys with rhenium deliver an increased open cell voltage and are also more preferred. It is surprising that palladium and iridium-additives start to be effective already at trace concentrations of about 10 ppm. This reduces the amount of expensive noble metals like iridium or palladium necessary for production of the electrocatalyst.

[0076] Preferred are palladium or iridium or rhenium contents between 0.1 and 50 at.-%. Since the open cell potential of the cathode decreases at palladium or iridium contents of over 20 at.-% palladium or iridium and in order to reduce costs contents while electrocatalytic activity rises between 10 ppm traces and 1 at.-%, contents between 1 at.-% and 10 at.-% are more preferred, contents between 1 at.-% and 5 at.-% are most preferred. Moreover alloys with lower contents such as 2 at.-% palladium or iridium have a better adhesion on nickel substrates. Ruthenium alloys comprising platinum are less effective than Ru-Pd- or Ru-Ir-alloys and are more expensive than alloys comprising palladium. For ruthenium alloys comprising rhenium (RuRe) rhenium contents between 1 at.-% and 10 at.-% are also more preferred as alloys with large rhenium contents are probably not resistant against the catholyte.

[0077] Ternary alloys of ruthenium $\text{RuMe}_I\text{Me}_{II}$ with a small amount of palladium (as Me_I) and iridium or rhenium (as Me_{II}) have even less polarization at current densities of $80\text{-}100 \text{ mA/cm}^2$.

Therefore ternary alloys RuPdIr with 1-5 at.-% palladium, 1-5 at.-% iridium, ruthenium balance are most preferred. Iridium or palladium may be replaced by rhenium.

[0078] Fig. 6 shows a polarization curve for a film of the alloy ruthenium-palladium-iridium (96 at.-% Ru, 2 at.-% Pd, 2 at.-% Ir) on nickel compared to a film of palladium-iridium (50 at.-% Ir) on nickel. Ternary ruthenium alloys $\text{RuMe}_I\text{Me}_{II}$ with a small amount of rhenium (1-5 at.-%) such as Ru-Pd-Re and Ru-Ir-Re are most preferred, too. Quaternary alloys Ru-Pd-Ir-Re with 1 at.-% - 5 at.-% rhenium deliver an advantageous more positive open cell potential and are optimum while other quaternary alloys $\text{RuMe}_I\text{Me}_{II}\text{Me}_{III}$ with $\text{Me}_{III} \neq \text{Re}$ offer little or no advantage.

[0079] The above electrocatalysts can be deposited directly on a current collector like carbon paper (e.g. Toray TGP-H-060) or used as supported catalysts on carbon, activated carbon or other high surface area carbons like carbon blacks commonly sold under the trademark "VULCAN XC72R" or "VULCAN XC-200" by the Cabot Corporation, Boston, MA. Carbon blacks usually require etching prior use in order to ensure wetting by electroless plating solutions. Etching may be done by nitric acid or other methods known for oxidizing carbon such as a mixture of nitric acid and sulfuric acid although other solutions comprising potassium permanganate (Hummers-Offeman process), and solutions for oxidation of graphite comprising potassium chlorate (Brodie or Staudenmaier process) may be also used, too. Alternatively carbon nanotubes, graphite nanotubes, doped polyaniline nanofibers or doped polyaniline nanotubes or other nanostructured materials can be used as support. Carbon nanotubes and graphite nanotubes are available from a large number of suppliers like Bayer MaterialScience AG, D-51368 Leverkusen, Germany. A catalyst loading of 5-50% ruthenium or ruthenium alloy is preferred, a loading of 5-30 % is most preferred.

[0080] For production of quaternary electrocatalysts comprising rhenium galvanic deposition of ruthenium-rhenium alloys or thermal decomposition of ammonium hexachlororuthenate comprising ammonium hexachloropalladate, ammonium hexachloroiridate and ammonium perrhenate in hydrogen may be used.

[0081] Polyaniline (PANI) nanotubes or nanofibers can be electrodeposited on an inert electrode or produced by polymerization of aniline using a template or in a solution of a surfactant like 0.05 M sodium dodecyl sulfate in an acid such as 1 M hydrochloric acid or perchloric acid. For example 10 ml solution of sodium dodecylsulfate in 1 M HCl is ultrasonicated for 1 hour. Aniline is added up to a concentration of 0.2 M and 10 ml 0.125 M solution of ammonium peroxodisulfate is added slowly to the stirred aniline solution kept at 0-5°C for 3 hours. The solution is filtered and rinsed several times with methanol and distilled water and dried and yields polyaniline nanotubes.

[0082] Polyaniline nanotubes can also be synthesized by interfacial polymerization at the

interface of two phases or polymerization through a semipermeable membrane. In interfacial polymerization for example 0.3 ml aniline is dissolved in 10 ml chloroform. 0.18 g ammonium peroxodisulfate is dissolved in an aqueous solution of the acid used as dopant (e.g. 1 M perchloric acid). The organic phase is overlaid by the aqueous solution of ammonium peroxodisulfate. PANI nanotubes that form at the interface of the two phases and dissolve in the aqueous phase within 10 minutes at room temperature can be separated by filtration of the aqueous phase. Polypyrrole nanotubes, polythiophene nanotubes and poly(3,4-ethylenedioxythiophene) nanotubes can be produced in a similar way.

[0083] According to an embodiment of the current invention the catalyst is coated by thin ion-conducting ionomer layer copolymer of tetrafluoroethylene and a perfluorovinyl ether sulfonic acid commonly sold under the trademark “NAFION” by E.I. DuPont de Nemours and Company, Wilmington, Delaware or “FUMION” commonly sold by Fuma-Tech GmbH, D-66386 St. Ingbert, Germany. This ionomer acts as ion-conductor layer that ensures ion-conducting contact of all catalyst particles to the electrochemical cell as well as transport of the oxidants hydrogen peroxide or oxygen and the product of the cathode reaction water. Besides the ionomer polymer layer protects the catalyst layer against abrasion and ensures wetting of the electrode by an electrolyte for hydrogen peroxide cathodes. Preferred is coating using a copolymer of tetrafluoroethylene and a perfluorovinyl ether sulfonic acid such as a 5% solution of “NAFION” (eq. wt. 1100 u) in a mixture of lower aliphatic alcohols and 15-20% water or “FUMION FLNA-905” dispersion (eq. wt.: 900 u, 5% dispersion in 45% water 50% n-propanol).

[0084] According to another embodiment of the present invention the topology of oxygen cathodes shown in Fig. 3A, 3B, 3C, 3D and 3E is adapted due to the requirements of contact of the three phases instead of two phases for hydrogen peroxide cathodes. Fig. 3A shows a cross section of an oxygen electrode according to an embodiment of the invention that consists of coated carbon fibers (300) as part of the current collector. The carbon fibers (300) of the oxygen cathode that are shown in Fig. 3B are coated by a catalyst layer on an intrinsically conducting adhesive (302). The carbon fibres (300) are used as current collectors. Single fibers (300) can be coated as shown in Fig. 3B or bundles of carbon fibers (312) or woven carbon fibre fabric can be used as current collectors as shown in Fig. 3C.

[0085] Fig. 3D shows a side view of a carbon fiber electrode that can be used as oxygen cathode for fuel cells. The coated carbon fibers (300) are mounted on a current collector substrate (312) like a titanium current collector, carbon paper or conducting polymer-carbon fiber composite by means for attaching the carbon fibers to the substrate (312) such as bonding with a conducting adhesive, clamping the uncoated ends of the fibers in a holder on a metal substrate. The resulting cathodes can

be bonded to a PEM-membrane (314) of a PEM fuel cell using a "NAFION"-coating as shown in Fig. 3D.

[0086] A cathode consisting of coated carbon fibers shown in Fig. 3B or Fig. 3C can be also used as hydrogen peroxide cathode of a hydrogen peroxide fuel cell or semi fuel cell.

[0087] As current collector material (300 in Fig. 3A-3C) other high surface area materials such as woven carbon fiber fabric, metal meshes, hollow carbon tubes, porous carbon such as carbon aerogel, and metal foams like titanium sponge may be used.

[0088] The current collector material (300) is coated by conducting adhesive (302), a catalyst layer (306) and an ionomer coating (310) as shown in Fig. 3A.

[0089] As catalysts supported catalysts like platinum, platinum alloys (306) or ruthenium on activated carbon (308) or unsupported catalysts can be used. For oxygen cathodes prior art platinum alloys may be used as preferred electrocatalysts.

[0090] According to a further embodiment of the present invention electrocatalysts for oxygen cathodes selected from the group consisting of platinum, ruthenium, osmium, rhenium, palladium, rhodium, iridium, chromium, cobalt, nickel, manganese, vanadium, silver, titanium, tungsten, aluminum, tin, silicon, molybdenum, bismuth, and alloys thereof are preferred.

[0091] The catalyst particles (304, 306) are bonded to the carbon fibers (300) by a conductive adhesive (302) comprising intrinsically electron conducting polymers.

[0092] The catalyst layer is coated by a thin ion-conducting ionomer layer (310) that ensures fast diffusion of oxygen that comprises a fluorinated ionomer such as "NAFION" or "FUMION".

"NAFION" can be also used to bond the fibers to a PEM membrane (314) of a fuel cell as an ion-conducting adhesive bridge as shown in Fig. 3D. The "NAFION" bridge can consist essentially of porous "NAFION" foam produced by suitable additives to the "NAFION" adhesive that produce gas during drying or heating such as azodicarboxamide, fluorinated hydrocarbons or carbon dioxide generating compositions comprising an acidic component and a alkali or earth alkali carbonate or alkali hydrogen carbonate. The pores of the "NAFION" foam improve diffusion of oxygen.

[0093] The topology of this oxygen cathode according to the present invention ensures fast diffusion of oxygen and good ionic conductivity as well as electrical conductivity.

[0094] Instead of coated carbon fibers hollow conducting tubes could be used for an oxygen electrode according to an embodiment of the present invention. The coating of the tubes with electrocatalyst may be bonded to the interior of the hollow conducting tubes. This could be accomplished by immersion of the tubes in a low viscosity electron conducting adhesive and coating the interior of the tubes with a fine catalyst powder by blowing with compressed air. After drying the tubes the electrocatalyst layer may be coated with a thin layer of an ion conducting

ionomer layer such as "NAFION".

[0095] Alternatively the hollow conducting tubes can be coated on the interior and exterior surface in order to increase the area of the electrode. A woven carbon fiber fabric may be coated in a similar way.

[0096] According to another embodiment of the present invention shown in Fig. 4A -Fig. 4D a porous current collector (400) like Titanium sponge can be coated in a similar way for use as an oxygen electrode. The substrate can be coated with conducting adhesive (402) by immersion, the electrocatalyst layer (406) can be applied by blowing catalyst powder on the adhesive using compressed gas. Coating with a ionomer layer like "NAFION" (410) can be performed by immersion, again. The topology of this design offers improved electrical conductivity of the adhesive that bonds the catalyst as well as fast diffusion of oxygen and hydronium ions (H_3O^+).

[0097] Cathodes of one of the previously described topologies using an intrinsically electron conducting adhesive in order to bond an electrocatalyst may be also employed for other fuel cell cathodes such as nitric acid cathodes, nitrous oxide (N_2O) cathodes, chlorate cathodes, chlorine cathodes or bromine cathodes. Furthermore the principle may be also applied to fuel cell anodes by using a reductively doped conducting polymer such as polyparaphenylene (PPP) or polythiophenes as intrinsically conducting adhesive for bonding of the anode electrocatalyst.

[0098] According to another embodiment of the present invention the previously described hydrogen peroxide cathodes or oxygen cathodes can be employed in a fuel cell using magnesium anodes, aluminium anodes or zinc anodes (502) as shown in Fig. 5.

[0099] Alternatively borohydride anodes, methanol anodes, formate anodes or formaldehyde anodes (502) can be used. For borohydride anodes an electrocatalyst like a platinum group metal like palladium on carbon (such as "Vulcan XC72R") may be used while for anodes using organic fuels a platinum-ruthenium or ruthenium decorated platinum electrocatalyst can be employed although other electrocatalysts might be also suitable.

[0100] For fuel cells using hydrogen peroxide cathodes (500) as shown in Fig. 5 the catholyte is separated from the anolyte of each cell by a polymer electrolyte membrane (PEM) (504). Catholyte and oxidizer is supplied by pipes (506, 508), anolyte and fuel by tubes (510). Cathodes (500) according to the present invention are mounted in an electrode holder consisting of two metal sheets (516, 518) and screws (520). A metal bar (512, 514) provides electrical contact to the electrodes of each cell. Titanium is preferred material for this holder and screws for the cathode and the metal bar. The electrodes can be arranged as bipolar electrodes as shown in Fig. 11 comprising a cathode (with a conductive substrate (200), an intrinsically electron-conducting adhesive layer (202), a catalyst layer (207) and an ionomer layer (210)) and an anode (220).

[0101] According to a further embodiment of the invention hydrogen anodes (502) can be combined with hydrogen peroxide cathodes according to the present invention as shown in Fig. 5. For such hydrogen anodes palladium or platinum on carbon (such as "Vulcan XC72R") can be used as electrocatalysts.

[0102] The catholyte according to the present invention further comprises an acid such as sulfuric acid, perchloric acid, an alkali hydrogen sulfate, ammonium hydrogen sulfate, sulfonic acids or carboxylic acids such as acetic acid because the electrochemical potential of hydrogen peroxide in acidic solutions is considerable larger this increases the open cell voltage of the fuel cell.

[0103] For oxygen cathodes in an aqueous catholyte perchloric acid is preferred because adsorption of sulfate ions reduces the activity of the electrocatalyst by adsorption.

[0104] For hydrogen peroxide cathodes acids are preferred that form peroxy acids with hydrogen peroxide. For example sulfuric acid instantaneously reacts with hydrogen peroxide to small amounts of Caro's Acid H_2SO_5 according to $\text{H}_2\text{SO}_4 + \text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{SO}_5 + \text{H}_2\text{O}$.

Although the equilibrium constant of this reaction is small ($K=3.125$) the small amount of H_2SO_5 of the order of 10 mM/l formed strongly influences the polarization of the cathode as can be shown in a comparison with a catholyte comprising only perchloric acid in Fig. 9.

When sulfuric acid is added the polarization of the cathode at high current densities is reduced compared to an electrolyte without sulfuric acid. A similar effect can be observed with acetic acid by formation of peracetic acid. For platinum electrocatalysts the open cell voltage is also increased in presence of Caro's acid.

[0105] A concentration of $c(\text{H}_2\text{SO}_4)$ of ≥ 0.5 mole/l and a concentration of $c(\text{H}_2\text{O}_2) \geq 1$ mole/l in the catholyte are preferred, a concentration of $c(\text{H}_2\text{O}_2) \geq 2.3$ M and $c(\text{H}_2\text{SO}_4) \geq 0.5$ mole/l is more preferred in order to produce a sufficient concentration of Caro's acid (H_2SO_5) in the catholyte. In this way polarization of hydrogen peroxide cathodes at high current densities is reduced by choice of the acid and concentration of acid and hydrogen peroxide.

[0106] Sulfuric acid, alkali hydrogen sulfates and carboxylic acids such as acetic acid, malonic acid, benzoic acid or phthalic acid are also preferred acids because the anions of those acids are not strongly adsorbed by the catalyst surface and do not hinder electrocatalytic activity. Moreover these acids form percarboxylic acids. Sulfuric acid or acetic acid are more preferred.

[0107] In dry state "NAFION" has a low proton conductivity. Therefore hydrogen peroxide fuel cells using a PEM membrane have to be stored filled with aqueous solutions. Alternatively hydrogen peroxide fuel cells that employ polymer electrolyte membranes (PEM membranes) for

separation of catholyte and anolyte can be stored in dry state without electrolyte if the polymer electrolyte membranes are wetted by a solution comprising an ionic liquid in order to increase the conductivity of the PEM membrane after storage. Because of the very low vapor pressure the ionic liquid is confined in the membrane. In an embodiment of the invention the solution is further comprising water in order to increase conductivity. 1-ethyl-3-methylimidazolium nitrate or 1-ethyl-3-methylimidazolium chloride can be used as water-miscible ionic liquid for wetting the PEM membranes during storage. Other ionic liquids such as 1-butyl 3-methyl imidazolium trifluoromethane sulfonate, 1-methyl-3-octyl-imidazolium trifluoromethanesulfonate, tetraalkylammonium nitrate or choline chloride-urea or ammonium nitrate-urea-acetamide might be also suitable for wetting of PEM membranes during storage.

[0108] Fuel cells using an oxygen cathode (300) according to an embodiment of the present invention have a membrane electrode assembly (MEA) as shown in Fig. 3D for PEM hydrogen-oxygen fuel cells. A PEM membrane (314) is used as electrolyte while the hydrogen anode is mounted on the other side of the PEM membrane. Hydrogen is fed through a diffusion layer to the anode coated with an electrocatalyst. Instead of the version of the invention of Fig. 3D a usual membrane electrode assembly (MEA) can comprise a cathode according to the present invention (comprising a gas diffusion layer as conductive substrate (200), an intrinsically electron-conducting adhesive (202), and an electrocatalyst layer (207)), a PEM-Membrane as Ionomer layer (210), an anodic electrocatalyst layer (230) and an anodic gas diffusion layer (232).

[0109] Fuel cells using an oxygen cathode according to an embodiment of the present invention and a borohydride anode, a methanol anode, a formaldehyde anode or a formate anode can use a fuel cell assembly shown in Fig. 3D, too. The anode is coated by an electrocatalyst like ruthenium-decorated platinum for methanol or other organic fuels or palladium for borohydride. The anode is wetted by an anolyte that contains the corresponding fuel. A PEM membrane (314) separates the cathode and the anolyte. For a fuel cell using an active metal anode such as a magnesium anode, an aluminium anode or a zinc anode the electrocatalytic anode is replaced by the active metal anode in a suitable anolyte.

[0110] As mentioned ruthenium and ruthenium-based alloys are superior electrocatalysts for fuel cell cathodes using concentrated hydrogen peroxide as oxidant. Such electrocatalyst layers may also deposited on inert substrates. Inert Substrates according to the present invention are resistant against the catholyte comprising the hydrogen peroxide oxidant such as carbon paper, carbon fiber fabric, activated carbon or carbon nanotubes bonded to a current collector. Nevertheless other materials such as conducting polymers like PANI or conducting polymer nanotubes may be used as substrate.

[0111] A ruthenium or ruthenium alloy electrocatalyst coating process according to the present invention comprises steps of pre-treating a provided substrate and coating the pretreated substrate. The step of pre-treating comprises cleaning the substrate in hydrochloric acid and distilled water. In an embodiment of the invention the pre-treating step is further comprising deposition of a single atom layer of palladium atoms as a catalyst for electroless deposition of the electrocatalyst.

[0112] According to an embodiment of the present invention the ruthenium or ruthenium alloy electrocatalyst can be deposited by an electrodeposition process. A ruthenium plating bath that contains a ruthenium nitridochloro complex $K_3[Ru^{IV}_2NCl_8(H_2O)_2]$ or a ruthenium nitrosyl complex is used to deposit ruthenium or ruthenium alloys. Preferred electroplating baths comprise a ruthenium nitridochloro complex. In a preferred embodiment the plating bath is further comprising sulfamic acid.

[0113] In a further embodiment of the present invention ruthenium or a ruthenium-based alloy can be deposited using an electroless plating bath on a support like activated carbon or high surface area carbon blacks or a substrate. An electroless plating bath comprising a ruthenium nitrosyl complex and a reducing agent such as dithionite and hydrazine or a ruthenium halide and an alkali borohydride can be used for this purpose. Pre-treating the substrate with a solution of palladium salt and a reducing agent such as tin(II)-chloride may be required for electroless plating of ruthenium and ruthenium alloys with the electroless plating baths comprising hydrazine on some substrates. Alternatively other reducing agents can be used.

[0114] Preferred electroless plating baths for production of electrocatalysts on a carbon black support comprise a ruthenium(III) chloride solution further comprising platinum metal halides. Sodium boranate solution is added dropwise at 5°C to the plating bath.

[0115] The preferred intrinsically electron-conducting adhesive according to an embodiment of the present invention comprises an intrinsically electron-conducting polymer such as polyanilines, polypyrroles, polythiophenes, polyparaphenylenes, polyazulenes, polyfuranes, polyindoles, polypyridines, polypyrazines, polytriazines, polythiazoles, polyimidazoles, polyquinolines, polybenzimidazoles, polytriazoles, polyoxydiazoles, polythianaphthenes, polycarbazoles, polybenzothiophenes, polybenzofuranes, polyheptadiyne, and polyparaphenylene vinylene and their substituted derivatives, copolymers, copolymers with alkenes or acrylates, and mixtures thereof. Polyaniline and PEDOT-PSS are more preferred. The preferred adhesive is further comprising a fluoropolymer ionomer as dopant and ion-conductor such as a copolymer of tetrafluoroethylene and a perfluorovinyl ether sulfonic acid commonly sold under the trademark "NAFION" by E.I. DuPont de Nemours and Company, Wilmington, Delaware, and a solvent such as dimethyl sulfoxide, N-methyl pyrrolidinone, dimethyl formamide, alcohols, water or a combination thereof.

[0116] Intrinsically electron conducting pressure sensitive adhesives (PSA) are more preferred as intrinsically electron-conducting adhesives because they are solvent-free adhesives that bond the electrocatalyst particles and connect them electrically to the current collector without wetting the electrochemically active upper side of the electrocatalyst particles. Preferred PSA-adhesives comprise a solution of polyaniline-dodecylbenzene sulfonic acid complex (PANI-DBSA), polystyrene-block-(polyethylene-ran-polybutylene)-block-polystyrene (SEBS) and a sufficient amount of poly- α -pinene as tackifier in p-xylene although other PANI-complexes like PANI-CSA, other polymers like polystyrene-block-polyisoprene-block-polystyrene or polystyrene-block-polybutadiene-block-polystyrene (SBS), or other tackifiers like rosin and rosin esters could be used, too. PSA adhesives can be easily applied for example by screen printing. This simplifies the manufacturing process of cathodes of the present invention.

[0117] Preferred fluoropolymer ionomer for adhesives comprising an ionomer is "NAFION" with an eq. wt. of 1100 u, preferred solvent is a mixture of dimethyl sulfoxide, alcohols and water. According to an embodiment of the present invention the adhesive can be further comprising graphite (for oxygen cathodes copper or silver may be used) and a further dopant such as dodecylbenzene sulfonic acid, camphor sulfonic acid or p-toluene sulfonic acid. Preferred additive is graphite powder.

BEST MODE OF CARRYING OUT THE INVENTION

[0118] The following examples illustrate the best mode of carrying out the embodiments of the invention. Examples 1-3 demonstrate the use of ruthenium coatings as electrocatalyst for hydrogen peroxide cathodes for fuel cells. Example 4 demonstrates measurement of a polarization curve of a massive ruthenium cathode.

EXAMPLES

Example 1

Preparation of a ruthenium electroplating bath

[0119] 1.97 g commercial ruthenium(III)chloride-hydrate ($\text{RuCl}_3 \cdot x\text{H}_2\text{O}$, reagent grade, 40.39% Ru, procured from Sigma-Aldrich, Taufkirchen, Germany) are dissolved in 78 ml deionized water. A solution of 11.625 g Sulfamic acid ($\text{NH}_2\text{SO}_3\text{H}$, p.a., $\geq 99\%$, procured from Fluka, Taufkirchen, Germany) in 78 ml deionized water is added and the solution is placed in a flask fitted with a Dimroth reflux condenser and the mixture is heated at the boil for 48 hours. During reflux the dark brown intransparent solution changes color to a transparent brown color. After cooling to room temperature the volume of the plating bath is adjusted to 310 ml (concentration about 30.6 mmole/l

Ru).

Example 2

Electroplating of a ruthenium layer on carbon paper

[0120] A 1 cm × 3 cm sheet of Toray TGP-H-120 carbon paper (procured from Quintech e.K., Goeppingen, Germany) is placed in a beaker filled with the ruthenium electroplating bath prepared according to example 1 that was heated prior use until the temperature of the bath reached 70°C. A 4 cm × 4 cm platinum sheet (procured from Oegussa GmbH, Vienna, Austria) is used as anode and ruthenium is deposited at a current density of 10 mA/cm² and a voltage of 2.5V for 2 minutes. After electroplating the carbon paper is rinsed with deionized water and dried. Under a microscope the deposited ruthenium coating is clearly visible.

Example 3

Electroplating of a smooth ruthenium layer on nickel for comparison of polarization curves

[0121] A 1 cm × 3 cm nickel sheet (99.9%, 0.1 mm thickness, procured from Alfa-Aesar GmbH&Co. KG, Karlsruhe, Germany) is used for ruthenium electroplating as described in example 2. A bright coating of ruthenium is deposited.

Example 4

Measurement of polarization curves for a massive ruthenium cathode

[0122] A ruthenium cathode is prepared from a 31.1 gram ruthenium ingot (99.95% Ru, Pt 205 ppm, Pd <1 ppm, Ir <1 ppm, Os 7 ppm, Rh 1 ppm, Ag <1 ppm, ACI Alloys Inc., San Jose, CA, USA). The polarization curve for the cathode of example 4 is nearly identical to the polarization curve of a thin ruthenium film according to example 3 in static solution of 2.32 M H₂O₂, 0.4 M H₂SO₄.

[0123] Examples 5-17 demonstrate electroplating of ruthenium-palladium-, ruthenium-iridium-, ruthenium-platinum-, ruthenium-rhodium- ruthenium-palladium-iridium-, ruthenium-rhenium-, ruthenium-palladium-rhenium-, ruthenium-iridium-rhenium-, and ruthenium-palladium-iridium-rhenium alloys for use as hydrogen peroxide cathode electrocatalysts. Examples 7, 8, 11, 12, and 13 demonstrate electroplating of most preferred electrocatalyst alloy films for comparison purposes.

Example 5

Preparation of a palladium electroplating bath

[0124] 0.275 g Palladium(II) chloride (PdCl_2 , procured from Riedel de Haen, Taufkirchen, Germany) are suspended in 16.9 g deionized water and 1.0 ml 25 % ammonia solution (pro analysi, procured from Fluka AG, Buchs, Switzerland) is added dropwise under stirring and heating at 70°C until all palladium chloride dissolves. 5 gram sulfamic acid (99%, procured from Fluka AG, Buchs, Switzerland) are added and the solution is filled into a flask equipped with a reflux condenser and heated for 24 hours at boiling temperature. After cooling 65.9 g deionized water are added to the solution to make up 82.5 ml palladium electroplating bath. The color of the solution changes to light yellow.

Example 6

Preparation of an iridium electroplating bath

[0125] 0.183 g Potassium hexachloroiridate(IV) (K_2IrCl_6) are dissolved in 16 ml deionized water. 2.30 g sulfamic acid are added and the mixture is boiled in a flask fitted with a reflux condenser for 48 hours. After cooling deionized water is added to make 70 ml electroplating bath.

Example 7

Preparation of a ruthenium-palladium-iridium electroplating bath

[0126] 8 ml of the ruthenium plating bath of example 1 are mixed with 0.266 ml palladium electroplating bath of example 5 and 0.245 ml iridium plating bath of example 6. Electroplating baths for ruthenium-palladium or ruthenium-iridium can be prepared by mixing above ruthenium plating bath with the above amounts of palladium or iridium electroplating baths.

Example 8

Electroplating of a ruthenium-palladium-iridium film on a nickel sheet

[0127] A $1\text{ cm} \times 3\text{ cm}$ nickel sheet (99%, procured from Alfa-Aesar GmbH&Co.KG, Karlsruhe, Germany) is placed in the electroplating bath of example 7 that was heated prior use until the temperature of the bath reached 70°C . A $4\text{ cm} \times 4\text{ cm}$ platinum sheet is used as anode and a ruthenium-palladium-iridium alloy is deposited at a current density of 11 mA/cm^2 for 1 minute. After electroplating the coated nickel sheet is rinsed with deionized water and dried. According to a preliminary analysis by SEM/EDX the deposited film consists of about 91.8 at.-% ruthenium, about 6.4 at.-% palladium and about 1.8 at.-% iridium.

Example 9

Measurement of a polarization curve for a hydrogen peroxide cathode according to example 8

[0128] A ruthenium-palladium-iridium coated nickel sheet prepared according to example 8 is used as a cathode in a fuel cell using a catholyte that contains 2.32 M H_2O_2 , 0.4 M H_2SO_4 against a magnesium anode (99.99% Mg, procured from Dead Sea Magnesium Ltd., Beer Sheva, Israel). The potential of the cathode is measured against a Palladium electrode that was in 0.5 M H_2SO_4 as reversible hydrogen electrode. Fig. 6 shows a polarization curve of of the ruthenium-palladium-iridium cathode in a static solution.

Example 10

Electroplating of a ruthenium-palladium-iridium film on a carbon paper

[0129] A 1 cm × 3 cm sheet of Toray TGP-H-120 carbon paper (procured from Quintech e.K., Goeppingen, Germany) is placed in a beaker filled with the electroplating bath prepared according example 7 that was heated prior use until the temperature of the bath reached 70°C. A 4 cm × 4 cm platinum sheet (procured from Oegussa GmbH, Vienna, Austria) is used as anode and ruthenium is deposited at a current density of 10 mA/cm² for 1 minutes. After electroplating the carbon paper is rinsed with deionized water and dried.

Example 11

Preparation of a rhenium electroplating bath

[0130] 0.27g of potassium perrhenate (KReO_4 , procured from Alfa-Aesar GmbH&Co. KG, Karlsruhe) are dissolved in 27 ml deionized water by stirring at 70°C. 0.8 ml conc. sulfuric acid (p.a., procured from Fluka AG, Buchs, Switzerland) and 0.675 g Magnesium sulfate ($\text{MgSO}_4 \cdot 7 \text{H}_2\text{O}$, p.a., procured from Fluka, Taufkirchen, Germany) are added.

Example 12

Preparation of a ruthenium-palladium-iridium-rhenium electroplating bath

[0131] According to example 7 an ruthenium-palladium-iridium-electroplating bath is prepared and 0.1 ml of the rhenium electroplating bath of example 11 are added.

Example 13

Electroplating of a ruthenium-palladium-iridium-rhenium film

[0132] Electroplating of a Ru-Pd-Ir-Re-alloy on a 1 cm × 3 cm nickel sheet is performed as in example 8.

Example 14

Preparation of a solution of $\text{Pt}(\text{NH}_3)_2(\text{NO}_2)_2$

[0133] 0.125 g potassium hexachloroplatinate(IV) (K_2PtCl_6 , procured from Alfa-Aesar GmbH&Co. KG, Karlsruhe, Germany) are suspended in 2 ml DI water. A concentrated solution of 1.250 g sodium nitrite (p.a., procured from Fluka AG, Buchs, Switzerland) in 2.83 ml deionized water are added. The mixture is heated to about 60°C under stirring for 30 minutes until all platinum salt dissolves. A pale yellow solution of $\text{K}_2\text{Pt}(\text{NO}_2)_4$ forms. When the solution has cooled to room temperature 50 μl 25% Ammonia solution (p.a., procured from Fluka AG, Buchs, Switzerland) are added.

Example 15

Preparation of a platinum electroplating bath

[0134] 0.806 g Sulfamic acid (p.a., procured from Fluka, Taufkirchen, Germany) are added to a solution of $\text{Pt}(\text{NH}_3)_2(\text{NO}_2)_2$ prepared according to example 14. 16.12 ml deionized water are added and the solution is heated at boiling temperature until a clear pale yellow solution is obtained.

Example 16

Preparation of a ruthenium-platinum electroplating bath

[0135] 8 ml of the ruthenium plating bath of example 1 are mixed with 0.2 ml platinum electroplating bath of example 15.

Example 17

Rhodium electroplating bath

[0136] 51.3 mg rhodium sulfate (procured from Sigma-Aldrich, Taufkirchen) are dissolved in 25.7 ml deionized water. 400 mg sulfamic acid are added. The yellow solution is heated at boil for 3 hours.

Examples 18-22 demonstrate the production of hydrogen peroxide (or oxygen) cathodes using most preferred supported or preferred unsupported electrocatalysts bonded by intrinsically conducting adhesives. The products and reaction mixtures of examples 20 and 22 should be handled with adequate safety precautions as occasional accidents (explosions) of by-products are reported in the literature.

Example 18

Preparation of an intrinsically electron-conducting adhesive

[0137] 29.5 mg Polyaniline (emeraldine base; PANI-EB $M_w=50,000$ g/mole, procured from

Sigma-Aldrich) were dissolved in 1.73 g Dimethyl sulfoxide (>99.5%, procured from Fluka AG, Buchs) under stirring at 60°C. After cooling 176.6 mg 5% "FUMION FL-905" (procured from Fuma-Tech GmbH, St. Ingbert) solution were added. 29.3 mg PANI-EB were dispersed in 0.1217 g of this PANI-EB-DMSO-ionomer solution. Instead of "FUMION FLNA-905" solution 5% "NAFION"-solution (procured from Sigma-Aldrich) in a mixture of alcohols and water may be used.

Example 19

Preparation of a fuel cell cathode using supported ruthenium electrocatalyst

[0138] Electron conducting adhesive prepared according to example 18 was applied to a 9.5×19 cm sheet of "Toray TGP-H-060" carbon paper (procured from Quintech e.K., Göppingen). 5% ruthenium on carbon electrocatalyst (procured from Alfa-Aesar GmbH&Co. KG, Karlsruhe, Germany) was dispersed on the adhesive layer and the adhesive was dried. The catalyst loading was 7.3 mg/cm^2 . After drying of the adhesive for 12 hours at room temperature and 5 minutes at 60°C a 5% dispersion of "FUMION FLNA-905" ionomer was applied to the surface of the electrode and the electrode was dried at room temperature.

Example 20

Preparation of a ruthenium-palladium-iridium black electrocatalyst

[0139] 314.5 mg Ruthenium(III) chloride ($\text{RuCl}_3 \cdot x \text{H}_2\text{O}$, Aldrich, Taufkirchen, Germany) were dissolved in 70 ml deionized water. A solution prepared by dissolving 5.8 mg Palladium(II) chloride (PdCl_2) in 3.42 ml deionized water and 0.1 ml 25% ammonia solution by stirring and heating and a solution of 14.1 mg Potassium hexachloroiridate(IV) (K_2IrCl_6 , procured from Alfa-Aesar GmbH&Co. KG, Karlsruhe) in 6.19 ml deionized water were added. The solution is cooled with an ice bath to +5°C.

[0140] 0.23 g Sodium borohydride (NaBH_4 , p.A., >96%, procured from Fluka, Taufkirchen) were dissolved in 9.04 ml deionized water and the solution was added dropwise by a dropping funnel under stirring within 30 minutes while the temperature of the ruthenium-palladium-iridium salt solution was kept between +6 and +8°C. Hydrogen evolved and ruthenium-palladium-iridium black forms. The solution was stirred for 12 hours, filtered through a sintered glass disc filter funnel (porosity G3) and the electrocatalyst was washed with deionized water, absolute ethanol and absolute ether.

Example 21

Preparation of a cathode with a ruthenium-palladium-iridium black catalyst

[0141] The electron conducting adhesive of example 16 was applied to a 1×3 cm sheet of “Toray TGP-H-060” carbon paper. Ruthenium-palladium-iridium-black electrocatalyst prepared in example 20 was dispersed on the adhesive layer and the adhesive was dried. The catalyst loading was 19.8 mg/cm^2 . After drying of the adhesive for 12 hours at room temperature and 5 minutes at 60°C a 5% dispersion of “FUMION FLNA-905” ionomer was applied to the surface of the electrode and the electrode was dried at room temperature.

Example 22

Measuring the polarization curve of the electrode according to example 21

[0142] The cathode electrode manufactured according to example 21 is placed in an holder that was manufactured from Poly(methylmethacrylate) (PMMA) commonly sold under the trademark “PLEXIGLAS” by Evonik Roehm GmbH, 64293 Darmstadt, Germany. The cathode is fixed by a strip of titanium sheet (procured from Small Parts Inc., Seattle, WA, USA) fastened to the holder by nylon or PTFE screws (procured from Small Parts Inc., Seattle, WA, USA). A 0.4 mm Haber-Luggin-capillary that consists of borosilicate glass commonly sold under the trademark DURAN by Schott AG Glaswerke, Mainz, Germany (now Duran-Group) is mounted in the holder about 0.8 mm in front of the cathode. A Pd-wire (procured from Aldrich, Taufkirchen) loaded with hydrogen by electrolysis prior use in 0.5 M H_2SO_4 (procured from Riedel de Haen, Taufkirchen) is used as reference electrode within the reference capillary. The cathode and the Haber-Luggin-capillary is placed in a static solution of 2.3 M H_2O_2 , 0.5 M H_2SO_4 (procured from Fluka, Taufkirchen/Buchs). A magnesium electrode (99.7%, procured from Fluka, Taufkirchen) was used as counter electrode. There is a considerable amount of oxygen generated by the ruthenium alloy black electrocatalyst of this electrode.

[0143] Example 22-28 demonstrate the manufacture of a most preferred ruthenium-palladium-iridium electrocatalyst on “Vulcan XC72R” carbon black and the most preferred intrinsically electron-conducting pressure sensitive adhesive. In a first step “Vulcan-XC72R” is etched by nitric acid in order to improve wettability. In Example 25 the manufacture of the most preferred intrinsically conducting pressure sensitive adhesive is demonstrated.

Example 23

Preparation of etched “Vulcan XC72R”

[0144] 0.2494 g “Vulcan XC72R” (Cabot Corporation, Boston, MA) were dispersed in 5.16 g concentrated nitric acid (pro analysi, 64-66%, procured from Fluka AG, Buchs, Switzerland) and heated to 65°C for 6 hours. The solution was cooled and filtered through a sintered glass disc filter funnel (G3 porosity) and washed with deionized water until the filtrate was neutral and the etched carbon black was dried.

Example 24

Preparation of a RuPdIr electrocatalyst on “Vulcan XC72R” (30% Ru load)

[0145] 165.5 mg etched “Vulcan XC72R” of example 23 were dispersed in 7.1965 g deionized water by stirring. 123.1 mg commercial ruthenium(III) chloride ($\text{RuCl}_3 \cdot x \text{H}_2\text{O}$, procured from Sigma-Aldrich, Taufkirchen, Germany) were dissolved in 34.94g deionized water. A palladium chloride solution (prepared by dissolving 5.5 mg anhydrous palladium(II) chloride (PdCl_2) procured from Riedel de Haen, Taufkirchen, Germany in 2.22 g deionized water and 0.3 ml 25% ammonia solution by stirring and heating to 70°C) and a solution of 17.3 mg Potassium hexachloroiridate(IV) (K_2IrCl_6 , procured from Alfa-Aesar GmbH&Co. KG, Karlsruhe, Germany) in 15.94 g deionized water were added. The solution is cooled with an ice bath to +5°C and stirred.

[0146] 0.1377 g Sodium borohydride (NaBH_4 , p.A., >96%, procured from Fluka, Taufkirchen, Germany) were dissolved in 3.7 g deionized water and the solution was added dropwise by a dropping funnel under stirring within 30 minutes while the temperature of the ruthenium-palladium-iridium salt solution was kept between +6 and +8°C. Hydrogen evolved and ruthenium-palladium-iridium deposits on the “Vulcan XC72R” carbon black. The solution was stirred for 12 hours, filtered through a sintered glass disc filter funnel (porosity G3) and the electrocatalyst was washed with deionized water, absolute ethanol and absolute ether and dried. The electrocatalyst is powdered using a mortar and pestle prior use.

Example 25

Preparation of an intrinsically conducting adhesive

[0147] 63.7 mg Polyaniline (emeraldine base; PANI-EB $M_w=50,000$ g/mol, procured from Sigma-Aldrich GmbH, Taufkirchen, Germany) were dissolved in 2.50 g Dimethyl sulfoxide (>99.5%, procured from Fluka AG, Buchs, Switzerland) under stirring at 60°C. The solution was cooled to room temperature. 253.5 mg of this PANI solution were placed in an aluminium dish and 92.4 mg of the above PANI-EB powder and 127.2 mg 5% “NAFION” solution (procured from Sigma-Aldrich GmbH, Taufkirchen) in lower alcohols and water were added.

Prospective example 26

Preparation of a conducting pressure sensitive adhesive

[0148] 60 mg PANI (emerald base) are placed in a beaker and 0.509 g Dodecylbenzene sulfonic acid (DBSA) are added. 5.5 ml absolute ethanol are added and the solution is heated for 2 hours under stirring at 50°C. Finally the ethanol is removed by distillation. The produced PANI-DBSA-salt is dissolved in 14 g p-xylene.

[0149] 6 g of this PANI-DBSA-solution in p-xylene are dissolved in 7 g xylene. 0.51 g polystyrene-block-(polyethylene-ran-polybutylene)-block-polystyrene (SEBS) copolymer are dissolved in 5.5 g p-xylene. 2.5 g poly- α -pinene are added. The PANI-DBSA-solution and the SEBS-poly- α -pinene PSA are mixed and yield an intrinsically conducting pressure sensitive adhesive.

Example 27

Preparation of a conducting pressure sensitive adhesive

[0150] 269 mg of colophony (procured from Fluka AG, Buchs, Switzerland are dissolved in 3.462 g 2-propanol (procured from Fluka, Taufkirchen, Germany). 72.5 mg of an aqueous solution of PEDOT:PSS (procured from Sigma-Aldrich, Taufkirchen) are added to 161.2 mg of the colophony solution. The prepared adhesive is applied to a carbon paper using a brush. The electrocatalyst is scattered on the adhesive layer while the adhesive is sticky.

Example 28

Preparation of a cathode with RuPdIr on "Vulcan XC72R"

[0151] A 3 cm×1 cm strip of Toray TGP-H-120 carbon paper is coated by the intrinsically electron conducting adhesive of example 25 or the intrinsically conducting pressure sensitive adhesive of example 26. The adhesive is dried prior coating the electrode with electrocatalyst if the pressure sensitive adhesive (PSA) of example 26 is used. Otherwise the wet adhesive layer is coated with electrocatalyst. 12.9 mg electrocatalyst prepared according to example 23 are distributed on the adhesive layer. If the PSA adhesive according to example 26 is used pressure can be applied to coat the electrode while dipping it in a pile of powdered electrocatalyst and an excess of electrocatalyst may be removed with a brush. Preferred is electrostatic application of the electrocatalyst powder. Otherwise the electrode is dipped into catalyst powder for coating.

[0152] The electrode is dried for 12 hours at room temperature and 2 minutes at 60°C. 1.0 ml 5%

"NAFION" solution in water/alcohol is dropped on the electrode. Excess solution may be removed from the edge of the electrode using a paper towel. The electrode is dried for 48 hours at room temperature prior use.

Example 29

Measurement of the polarization curve of a cathode prepared according to example 28

[0153] A cathode prepared according to example 27 is fixed in the holder according to example 22. As in example 22 a Haber-Luggin capillary is used for the RHE reference electrode in 0.5 M sulfuric acid. A magnesium electrode (99.7%) is used as anode. Fig. 8 shows the polarization curve obtained.

[0154] Examples 30-32 demonstrate the use of catholytes comprising Caro's-Acid or Peracetic acid.

Example 30

Catholyte comprising Caro's acid

[0155] 20.00 ml 35% hydrogen peroxide (pro analysi, $d=1.12924$ g/ml at 21.5°C , about 34.85% content after storage, procured from Fluka, Taufkirchen, Germany) are mixed with 29.376g deionized water in a volumetric flask. 5.1349 g concentrated sulfuric acid (98%, pro analysi, procured from Fluka AG, Buchs, Switzerland) are added and deionized water is added to make up to 100 ml volume. The solution contains about 0.513 mol/l H_2SO_4 , about 2.314 mol/l H_2O_2 and immediately forms about 0.009 M H_2SO_5 . Fig. 9 shows a polarization curve of a 1 cm× 3 cm ruthenium coated nickel sheet against a magnesium anode in this electrolyte.

Example 31

Catholyte comprising peracetic acid

[0156] 170 μl Acetic acid (>98%) are added to 14.834g of the catholyte of comparative example 32. The solution is about 0.2 M in acetic acid. The polarization curve of a 4 cm× 4 cm platinum sheet against a magnesium anode in this electrolyte is measured.

Comparative example 32

Catholyte comprising hydrogen peroxide in 1 M perchloric acid

[0157] 15 ml 1 M perchloric acid was prepared from 70% perchloric acid (p.a.). 12.37 g of this

solution was placed in a 25 ml volumetric flask. 5.00 ml 35% hydrogen peroxide (p.a., d=1.12924 g/ml at 21.5°C) are added and 1 M perchloric acid is added to make up 25 ml volume.

Fig. 9 shows the polarization curve of a 1 cm× 1 cm ruthenium coated nickel sheet against a magnesium anode in this electrolyte.

[0158] Although the present invention has been described in considerable detail with reference to certain preferred embodiments thereof, other versions are possible. For example it is possible to bond a carbon black layer as a pre-catalyst instead of the electrocatalyst on the adhesive layer that will be converted to the electrocatalyst by applying a metal salt and a reducing agent. Therefore, the spirit and scope of the appended claims should not be limited to the description of the preferred embodiments contained herein.

[0159] All the features disclosed in this specification (including any accompanying claims, abstract and drawings) may be replaced by alternative features serving the same, equivalent or similar purpose, unless expressly stated otherwise. Thus, unless expressly stated otherwise, each feature disclosed is one example only of a generic series of equivalent or similar features.

[0160] Insofar the description above and the accompanying drawings disclose any additional subject matter that is not within the scope of the claims below, the inventions are not dedicated to the public and the right to file one or more applications to claim such additional inventions is reserved.

CLAIMS:

What is claimed is:

1. A cathode for a fuel cell, which comprises
 - (a) a current collector;
 - (b) a layer of an intrinsically electron-conducting adhesive on a surface of said current collector;
 - (c) a layer of an electrocatalyst wherein dry particles consisting only of an electrocatalyst are directly bonded to the current collector by said electron-conducting adhesive, wherein said particles are selected from the group consisting of electrocatalyst powder, electrocatalytically coated nanotubes, and electrocatalytically coated nanofibers; and
 - (d) an ionomer layer in contact to said electrocatalyst,wherein said intrinsically electron-conducting adhesive is a pressure sensitive adhesive and wherein said intrinsically electron-conducting adhesive has a sufficient bonding strength for catalyst particles to ensure electrical contact between the electrocatalyst and the current collector.
2. A cathode according to claim 1, wherein said intrinsically electron-conducting adhesive comprises at least one intrinsically electron-conducting polymer that is selected from the group consisting of polyaniline, substituted polyanilines, polythiophene, poly(3,4-ethylenedioxythiophene), other substituted polythiophenes, polypyrroles, polyparaphenylenes, polyazulenes, polyfuranes, polyindoles, polypyridines, polypyrazines, polytriazines, polythiazoles, polyimidazoles, polyquinolines, polybenzimidazoles, polytriazoles, polyoxydiazoles, polythianaphthenes, polycarbazoles, polybenzothiophenes, polybenzofuranes, polyheptadiyne, polyparaphenylene vinylenes, and copolymers thereof.
3. A cathode according to claim 2, wherein said electron-conducting adhesive further comprises a dopant.
4. A cathode according to claim 3, wherein said dopant is an alkyl-substituted arylsulfonic acid with at least one alkyl-substituent with at least 3 carbon atoms.

5. A cathode according to claim 3, wherein said dopant is selected from the group consisting of dodecylbenzene sulfonic acid, dinonylnaphthalene sulfonic acid, and camphorsulfonic acid.
 6. A cathode according to claim 3, wherein said electron-conducting adhesive further comprises at least one tackifier.
 7. A cathode according to claim 6, wherein said electron-conducting adhesive further comprises an elastomer.
 8. A bipolar electrode comprising a cathode according to claim 1, further comprising an anode that is electrically connected to said cathode and mechanically supports said cathode.
 9. A membrane electrode assembly comprising a cathode according to claim 1, further comprising an ionomer membrane that is bonded to the ionomer layer of said cathode and an anode that is bonded to the other side of the ionomer membrane.
 10. A cathode according to claim 1, wherein said electrocatalyst comprises ruthenium or a ruthenium alloy.
-

11. A method of using an intrinsically electron-conducting adhesive for bonding electrocatalysts within a process of producing a membrane electrode assembly for a fuel cell comprising the steps of:
- (a) applying a layer of an intrinsically electron-conducting pressure sensitive adhesive on a surface of a conducting substrate for a cathode; and
 - (b) distributing a dry electrocatalyst powder on the layer of said electron-conducting pressure sensitive adhesive, wherein said electrocatalyst powder consists only of particles of a supported or an unsupported electrocatalyst and wherein said particles are selected from the group consisting of electrocatalyst powder, electrocatalytically coated nanotubes, and electrocatalytically coated nanofibers .
12. A method according to claim 11, wherein said intrinsically electron-conducting adhesive comprises:
- (aa) an intrinsically electron-conducting polymer that is selected from the group consisting of polyaniline, substituted polyanilines, polythiophene, poly(3,4-ethylenedioxythiophene), other substituted polythiophenes, polypyrroles, polyparaphenylenes, polyazulenes, polyfuranes, polyindoles, polypyridines, polypyrazines, polytriazines, polythiazoles, polyimidazoles, polyquinolines, polybenzimidazoles, polytriazoles, polyoxydiazoles, polycarbazoles, polybenzothiophenes, polybenzofuranes, polyheptadiyne, polyparaphenylene vinylene, and copolymers thereof;
 - (bb) a dopant; and
 - (cc) a tackifier.
13. A method according to claim 12, wherein said dopant is selected from the group consisting of dodecylbenzene sulfonic acid, dinonylnaphthalene sulfonic acid, and camphorsulfonic acid.
14. A method according to claim 12, wherein said dopant is an alkyl-substituted aryl sulfonic acid with at least one alkyl-substituent with at least 3 carbon atoms.

15. A fuel cell, comprising an anode, an electrolyte and a cathode comprising:
- (a) an electron-conducting current collector;
 - (b) an intrinsically electron-conducting adhesive on a surface of said current collector;
 - (c) a dry electrocatalyst powder that consists only of particles of an electrocatalyst which is directly bonded to the current collector by said electron-conducting adhesive; and
 - (d) an ionomer layer in contact to said electrocatalyst,
- wherein said intrinsically electron-conducting adhesive is a pressure sensitive adhesive and wherein said intrinsically electron-conducting adhesive has a sufficient bonding strength for catalyst particles to ensure electrical contact between the electrocatalyst and the current collector.
16. A fuel cell according to claim 15, wherein said electrocatalyst is an ruthenium alloy on carbon with at least 50 at.-% ruthenium further comprising at least 0.01 at.-% of at least two metals selected from the group consisting of palladium, iridium, rhenium, platinum, osmium, and rhodium.
17. Fuel cell according to claim 15, wherein said electrocatalyst is a ruthenium alloy on carbon further comprising at least 0.1 at.-% and up to 10 at.-% of at least two metals selected from the group consisting of palladium, iridium, rhenium, and platinum.
18. A fuel cell according to claim 15, wherein said electrocatalyst is a supported catalyst with a ruthenium alloy comprising 1 at.-% to 10 at.-% based palladium and 1 at.-% to 25 at.-% iridium.

Fig. 1 Prior art cathode

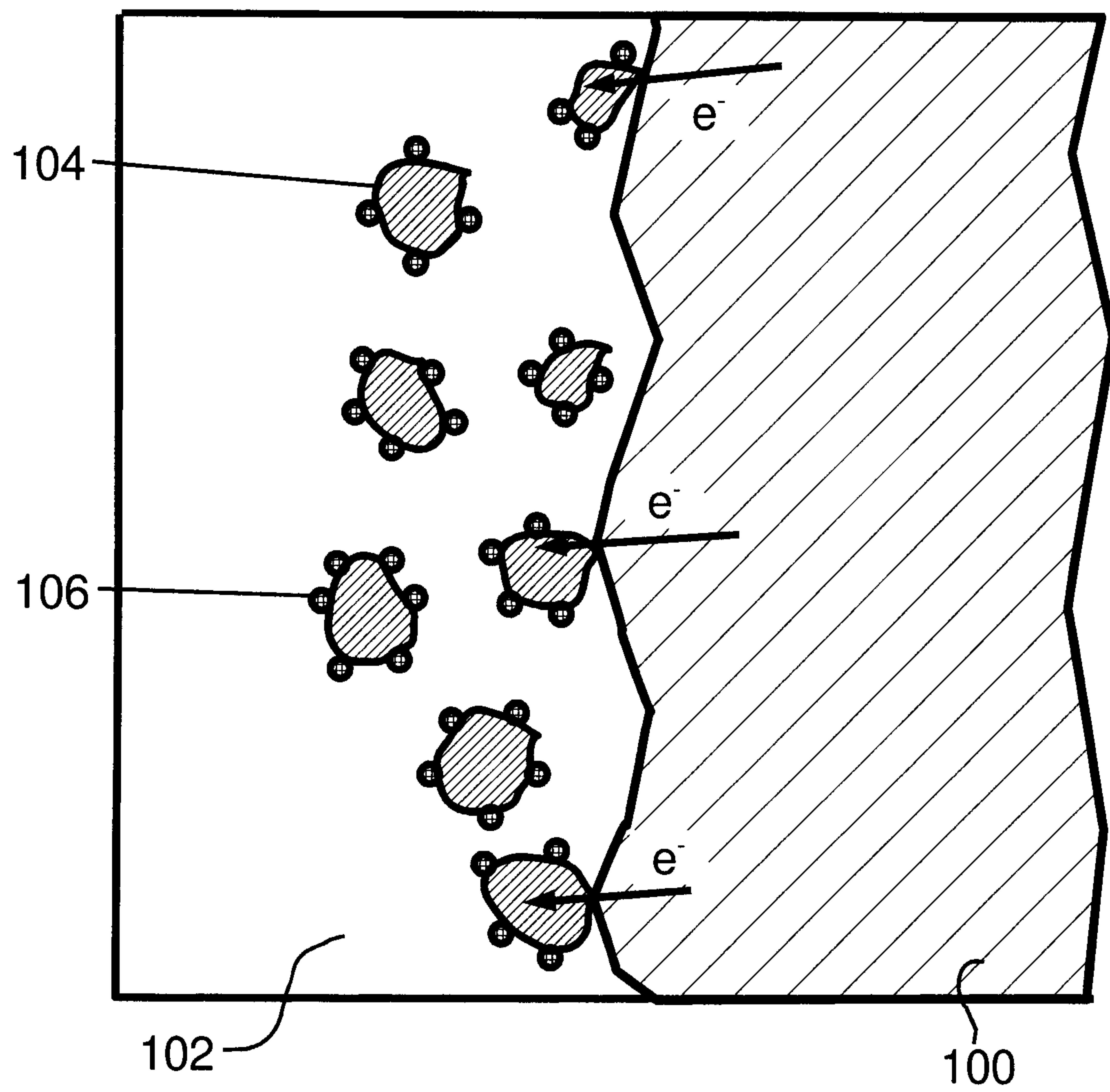


Fig. 2A

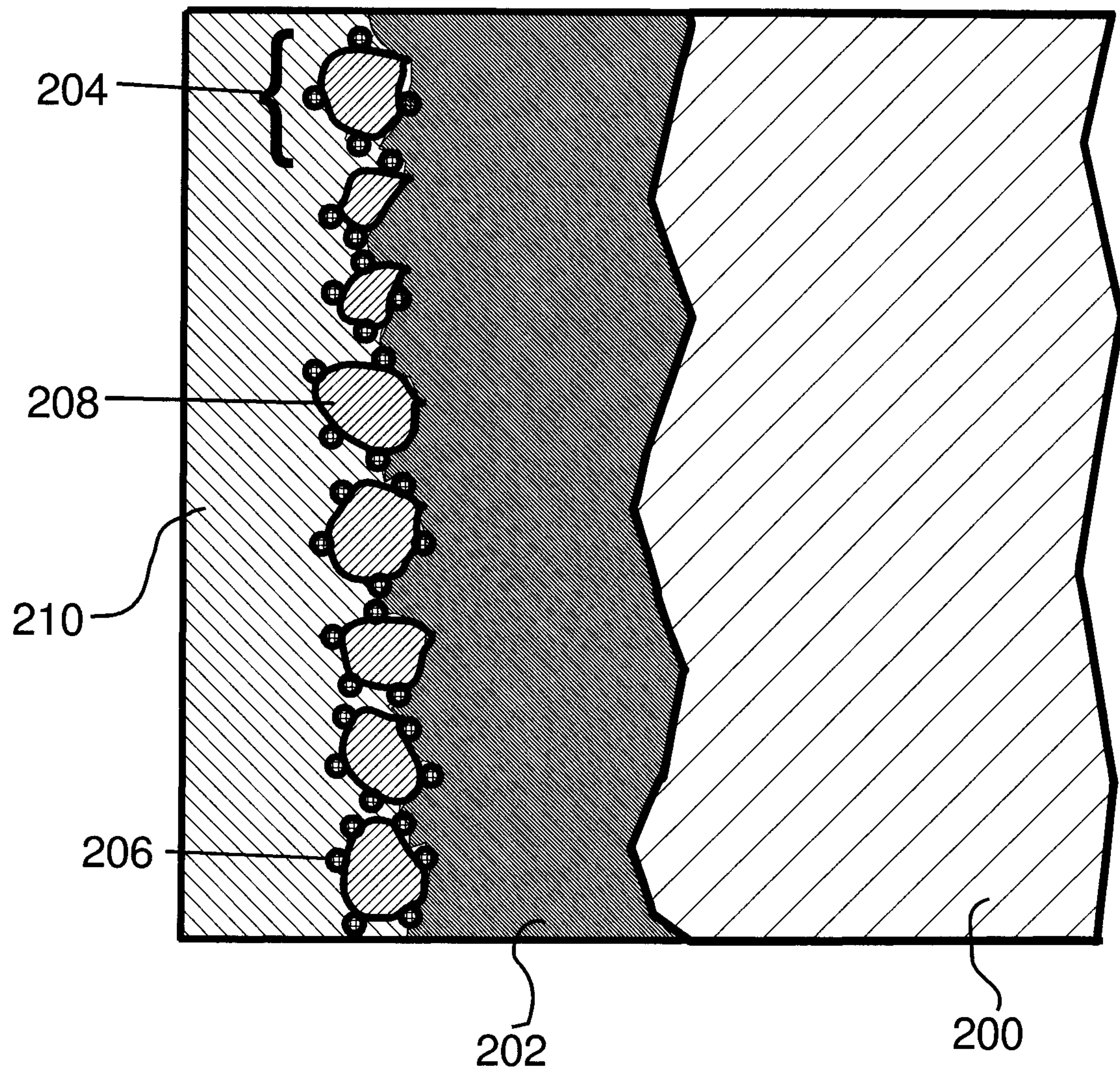


Fig. 2B

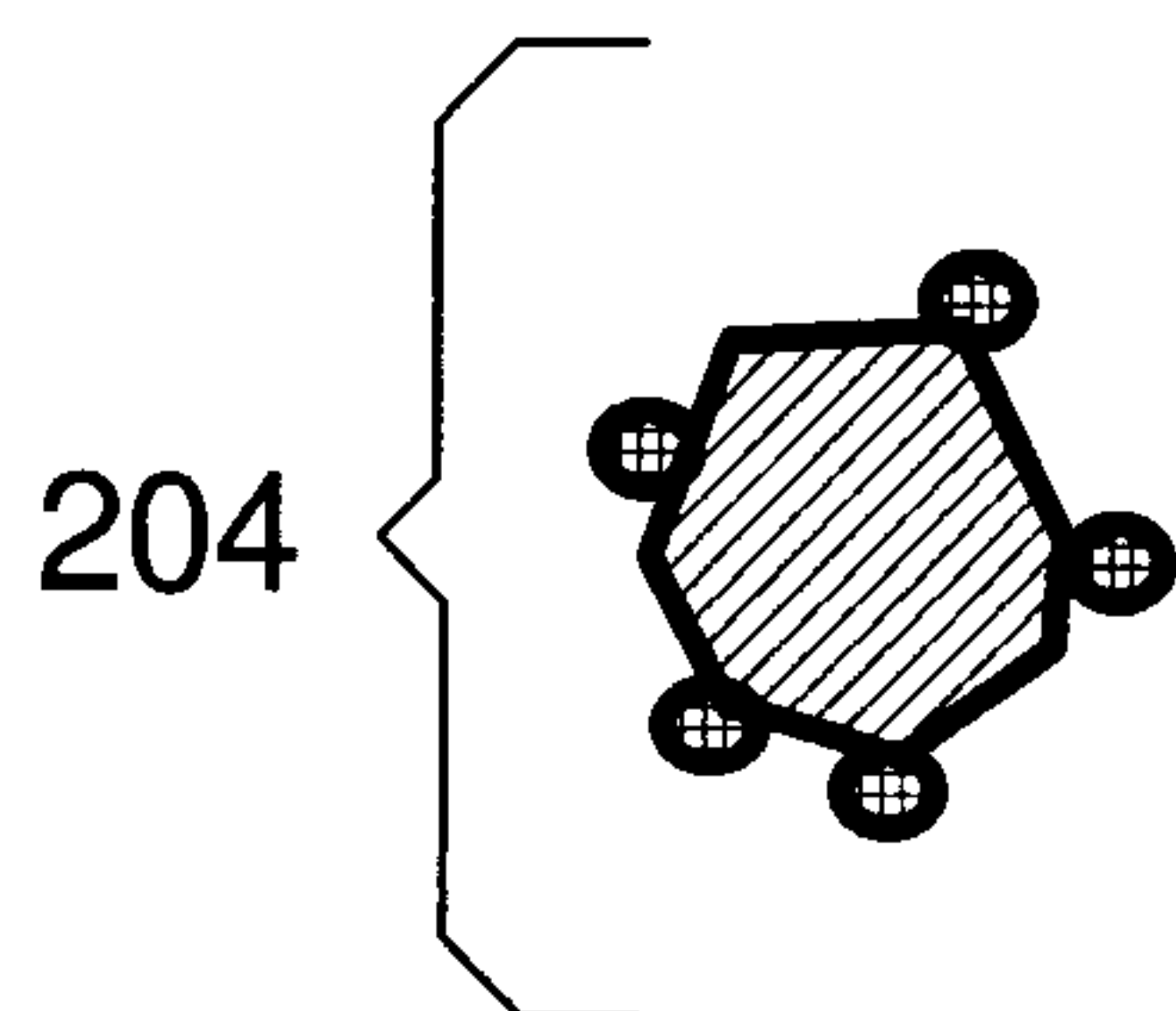


Fig. 2C

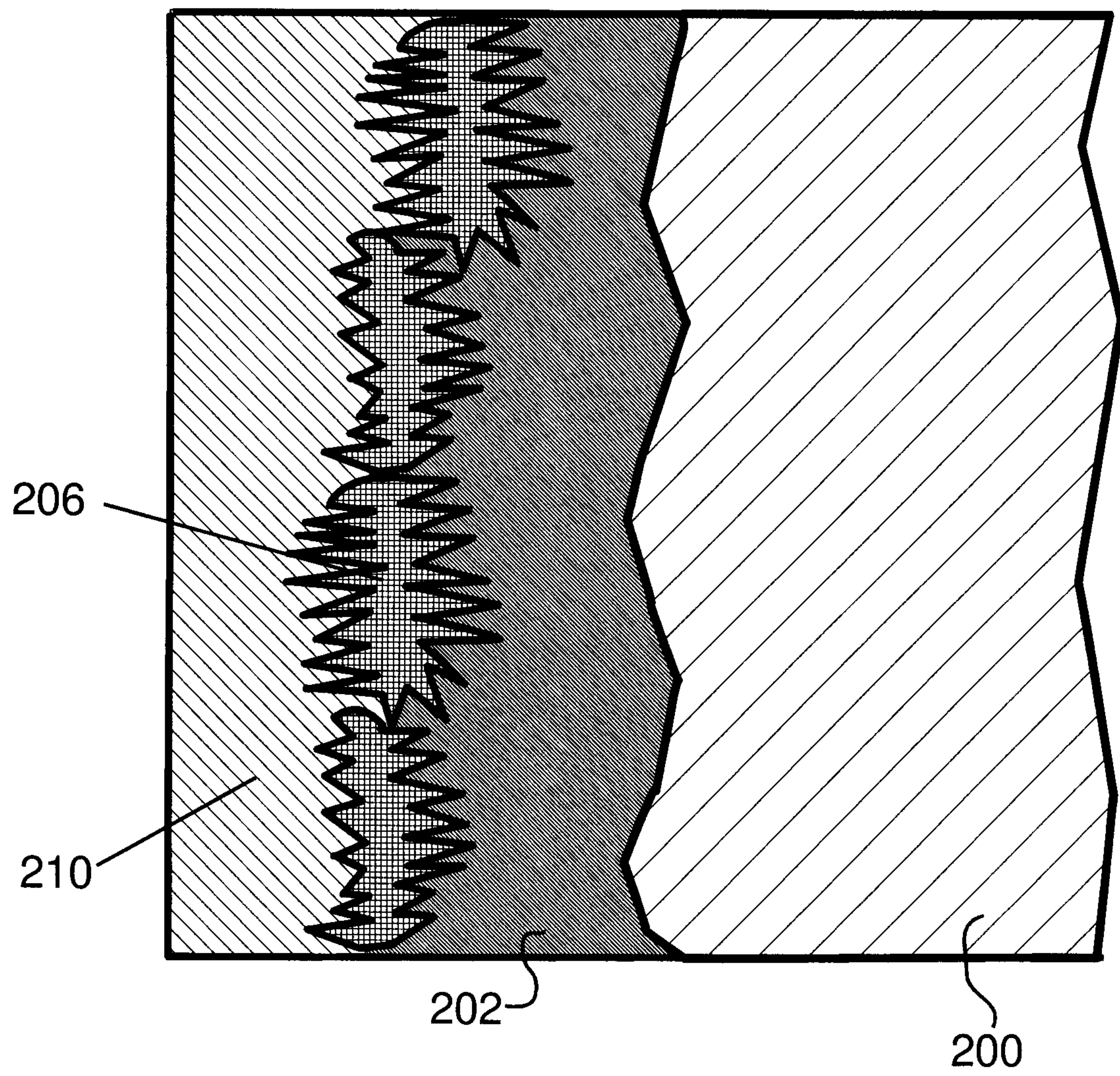


Fig. 2D

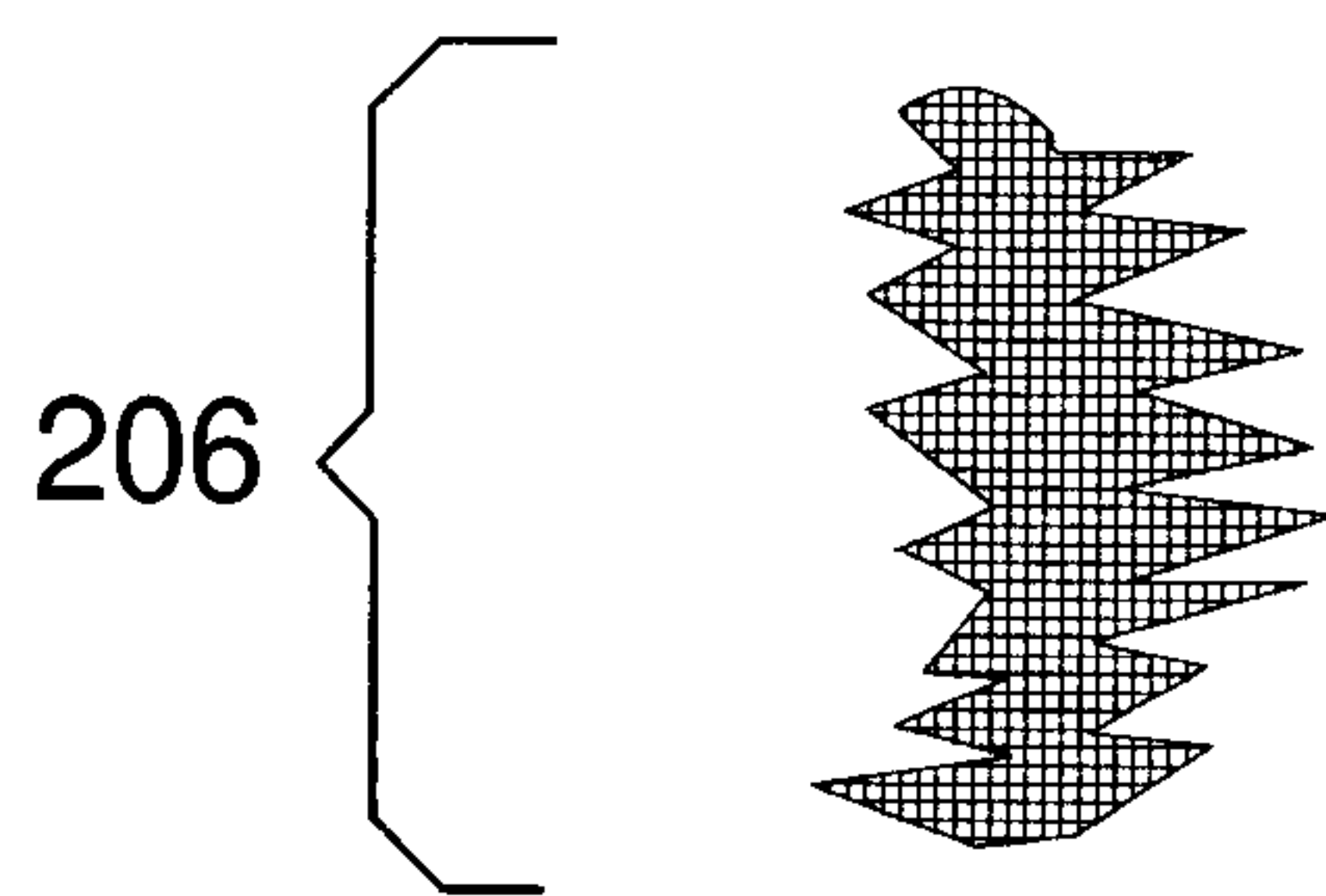


Fig. 3A

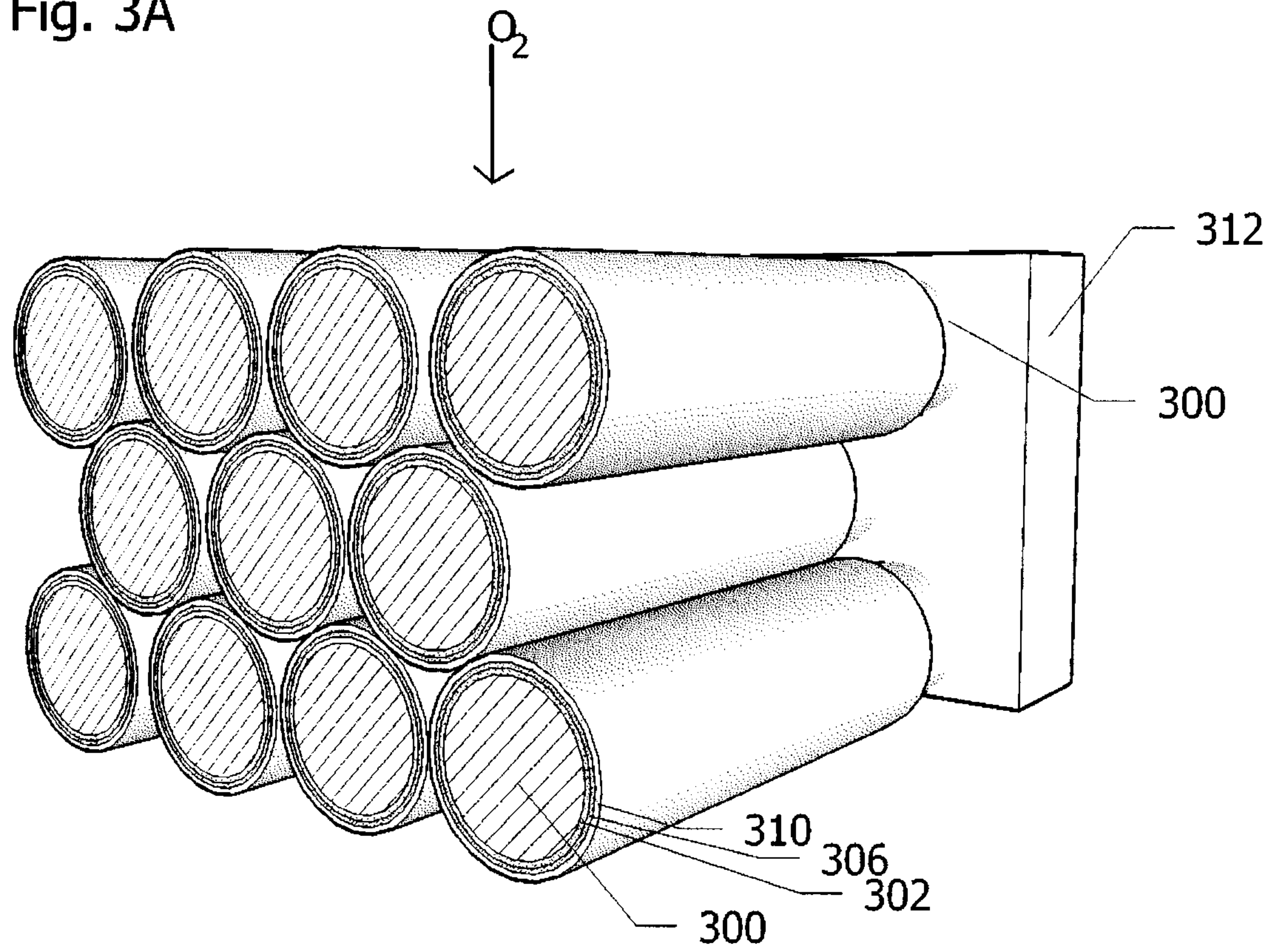


Fig. 3B

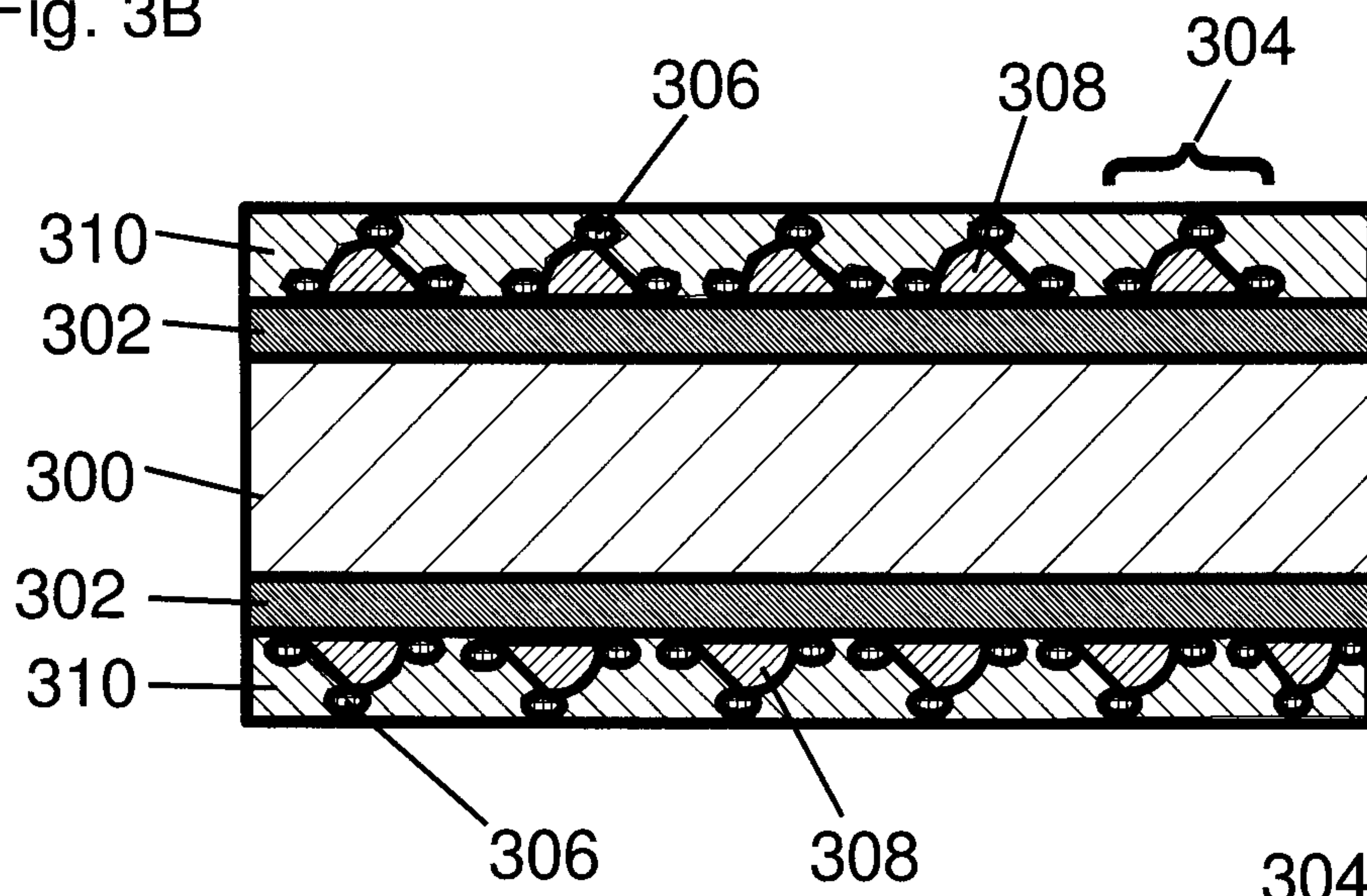


Fig.3C

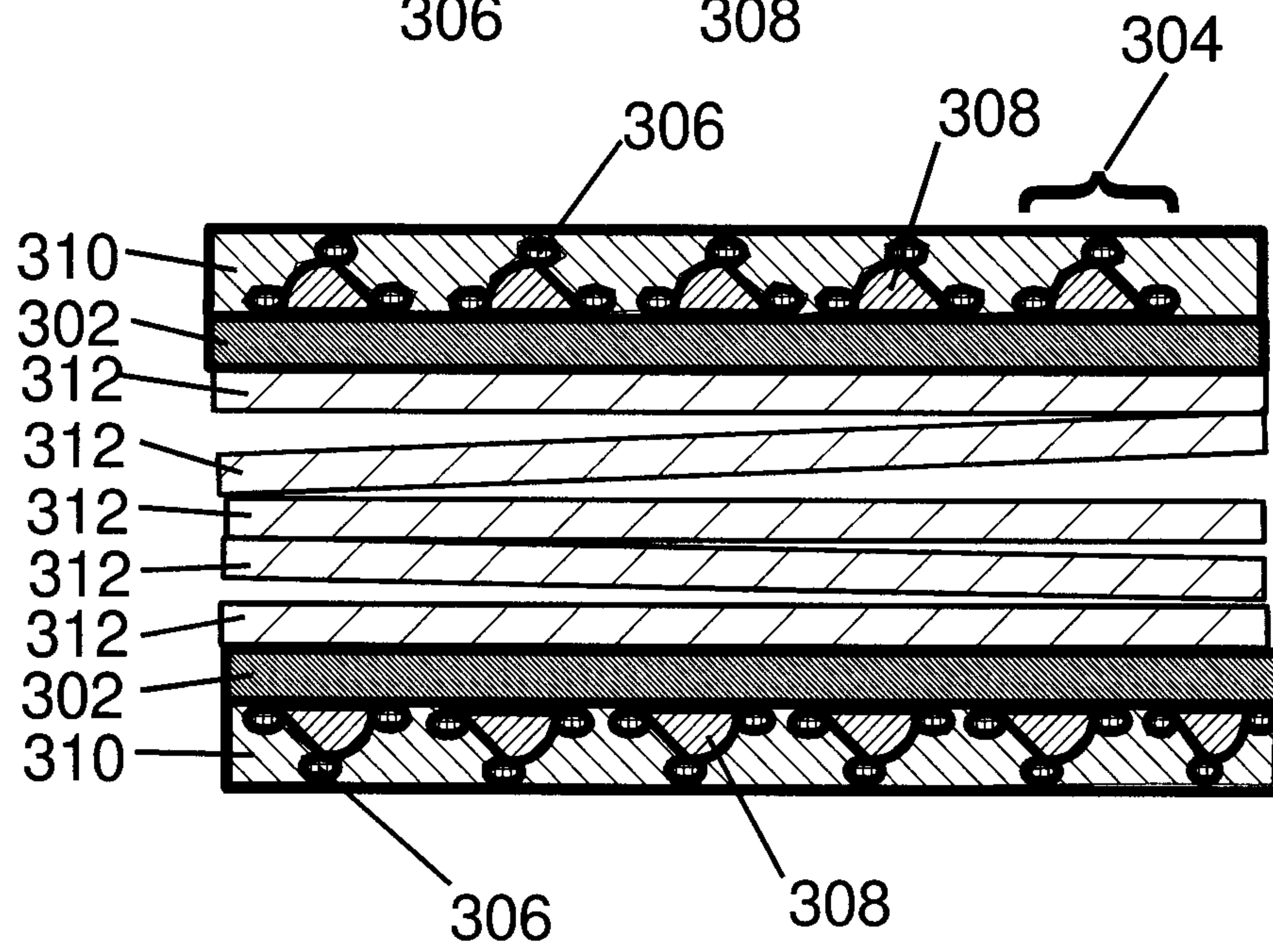


Fig. 3D

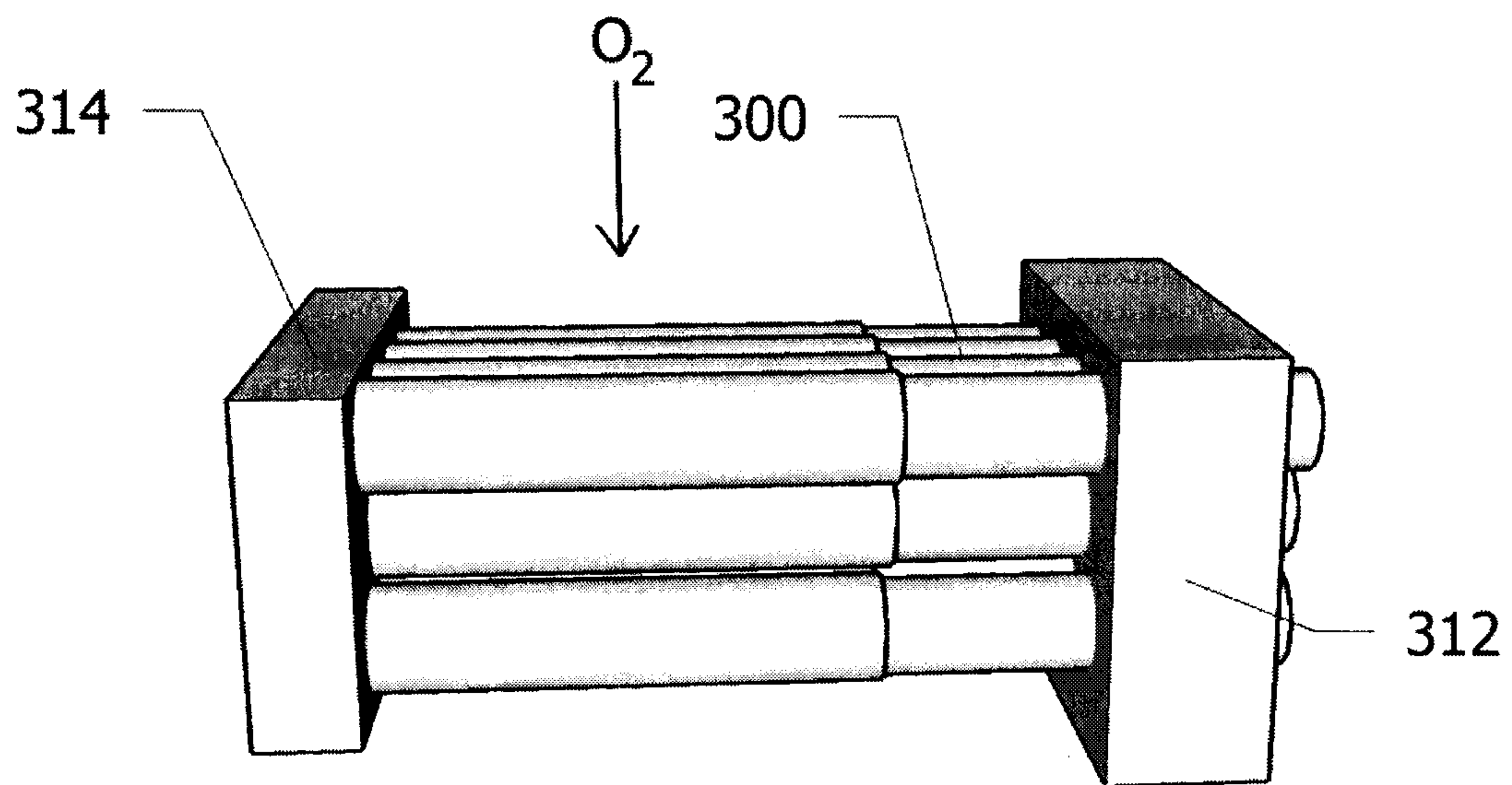


Fig. 3E

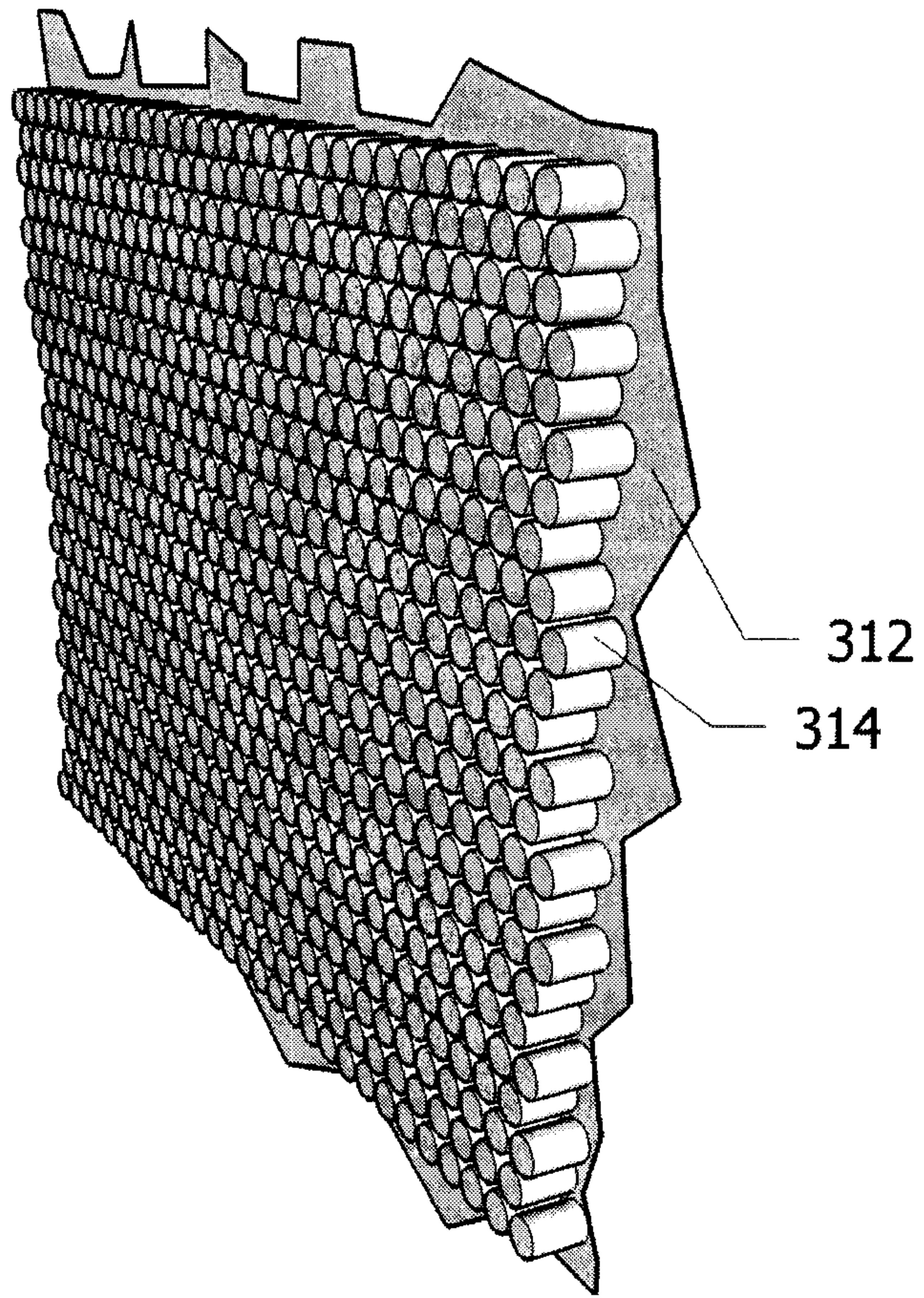


Fig. 4A

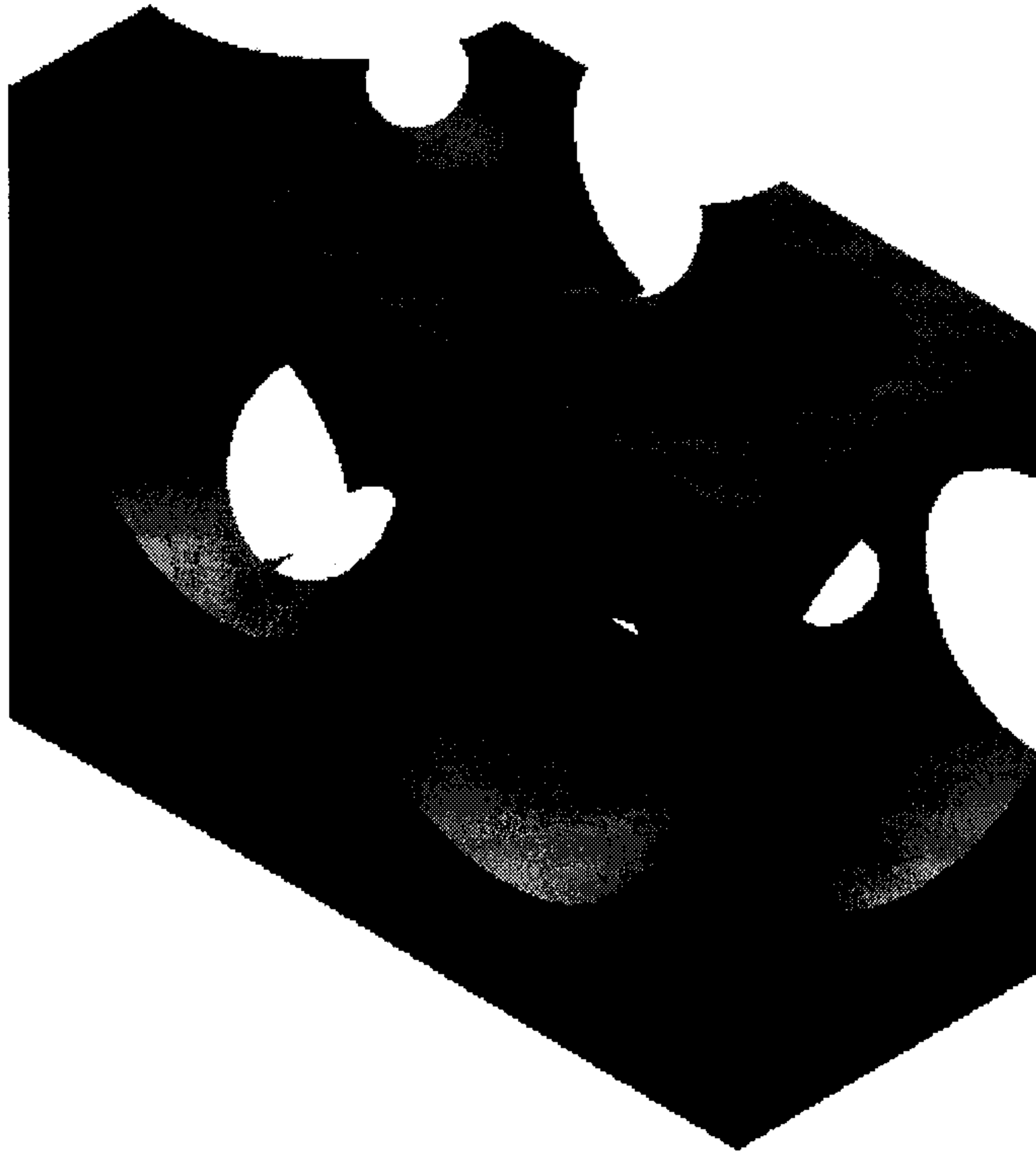


Fig. 4B

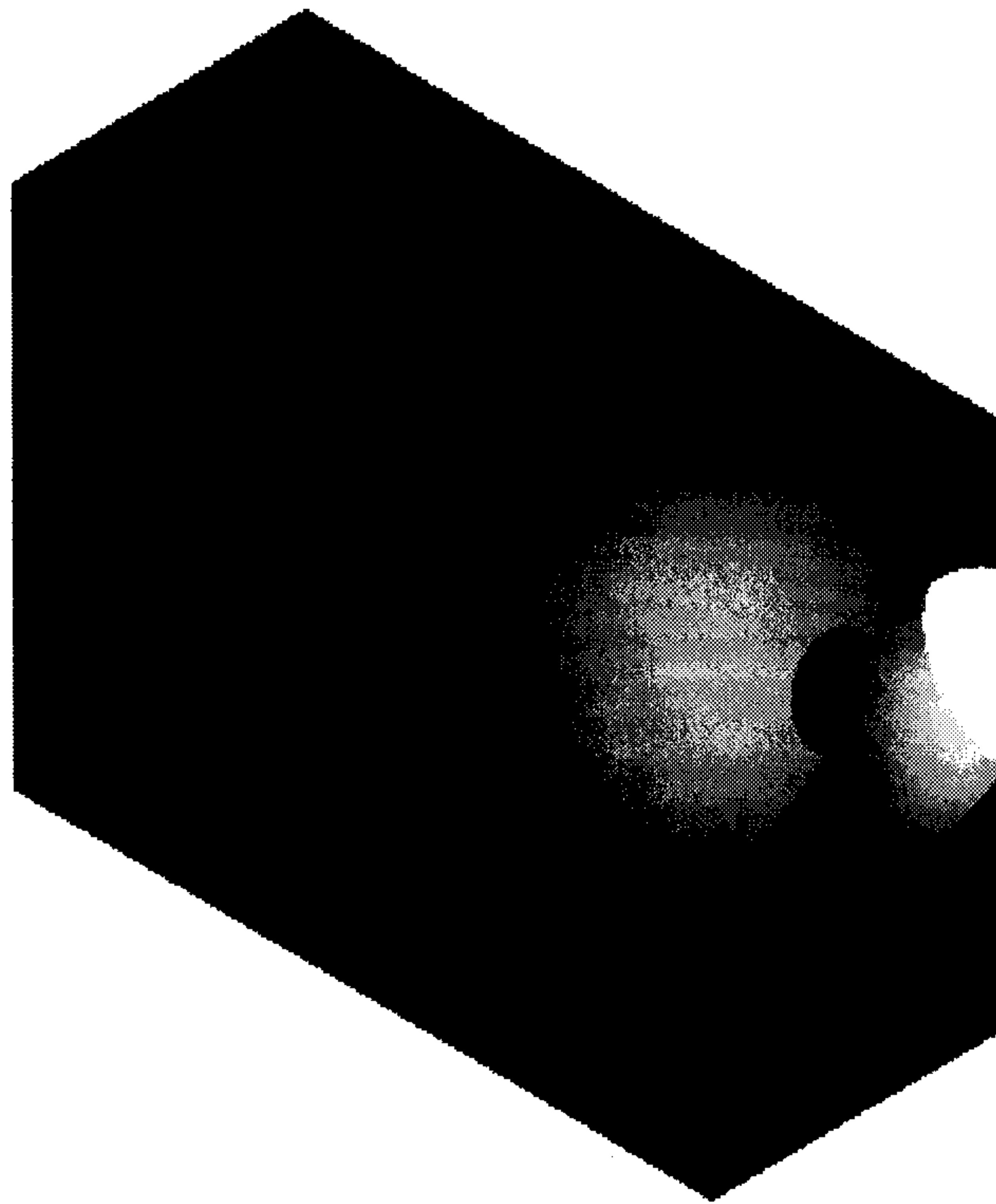


Fig. 4C

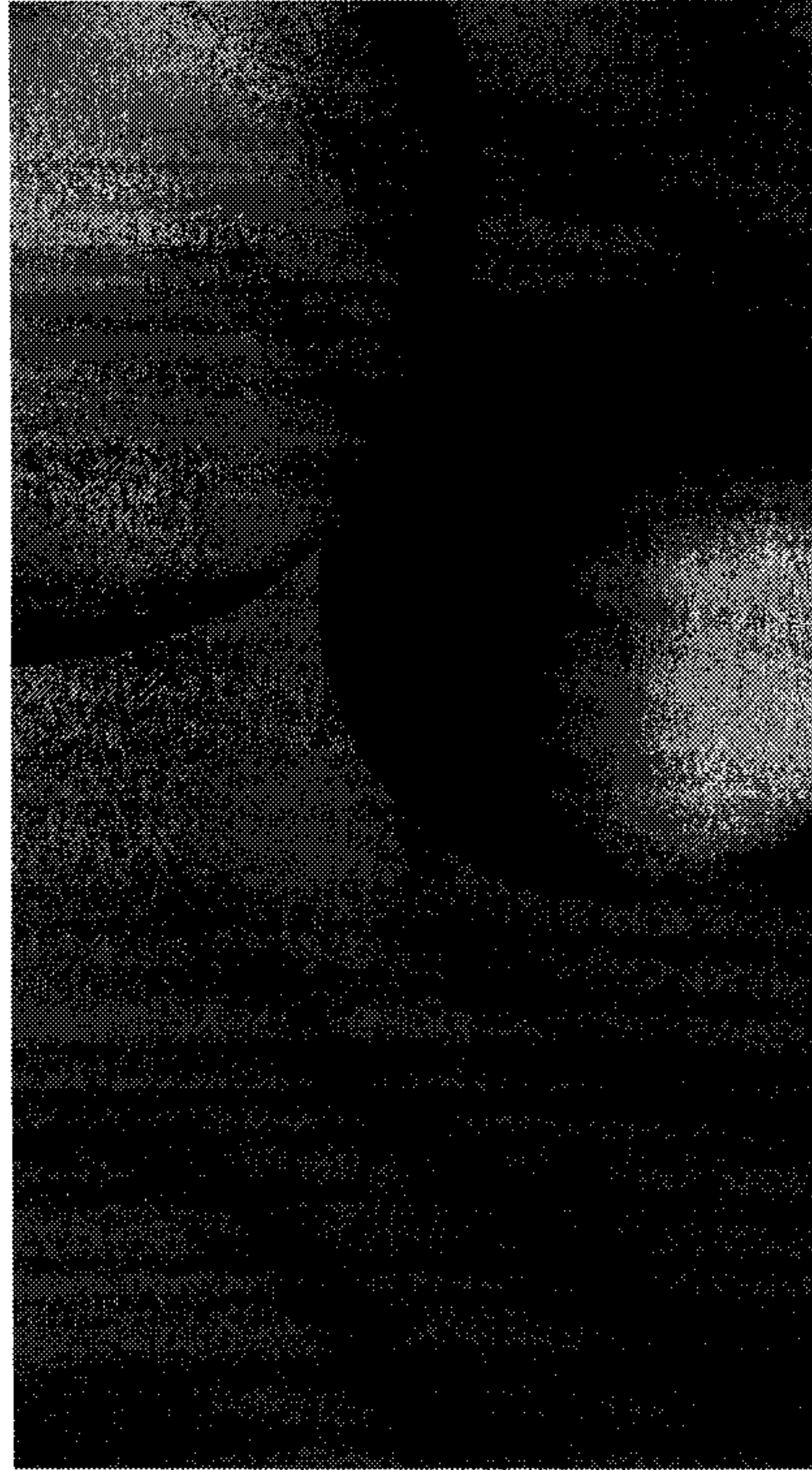
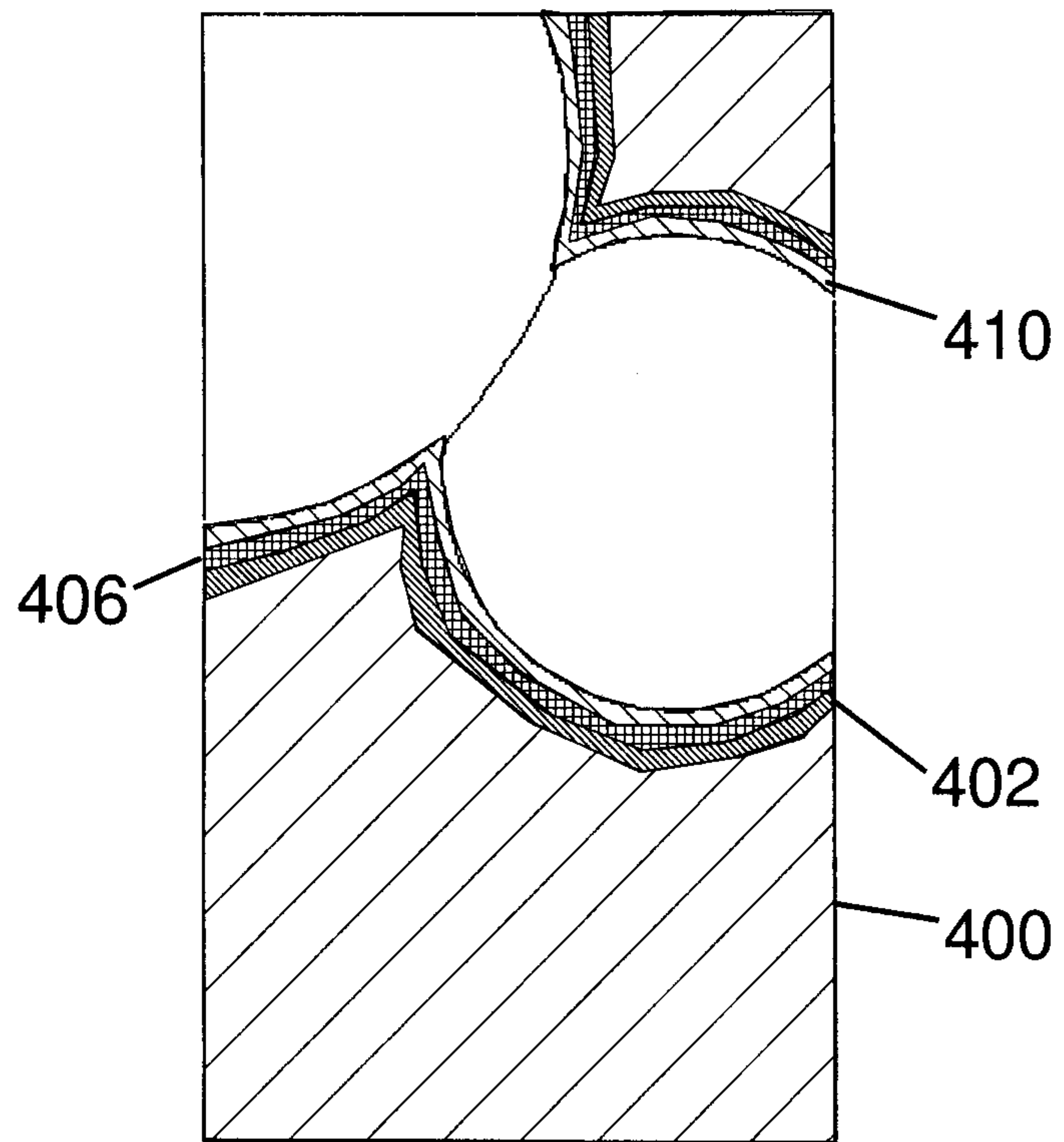


Fig. 4D



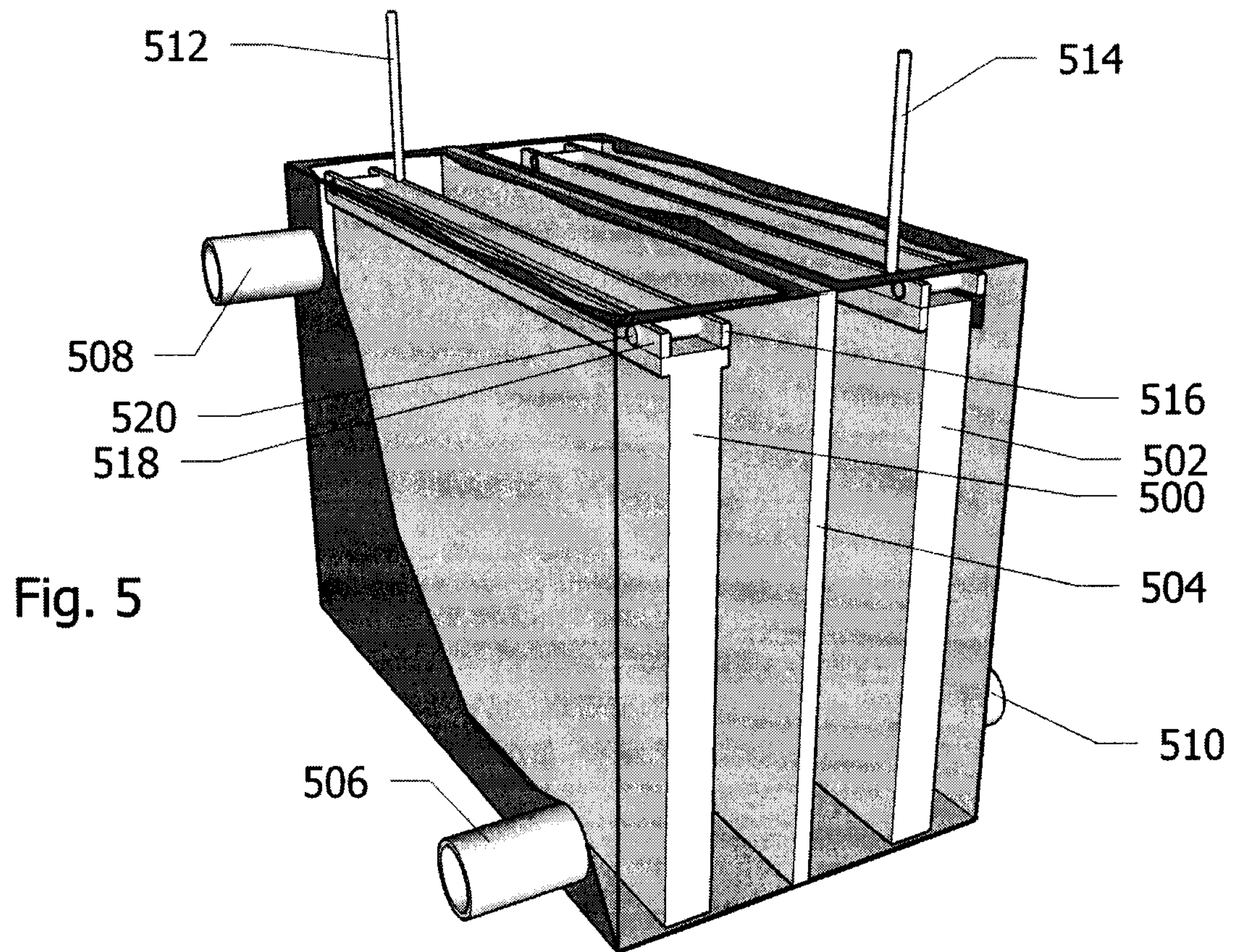


Fig. 6

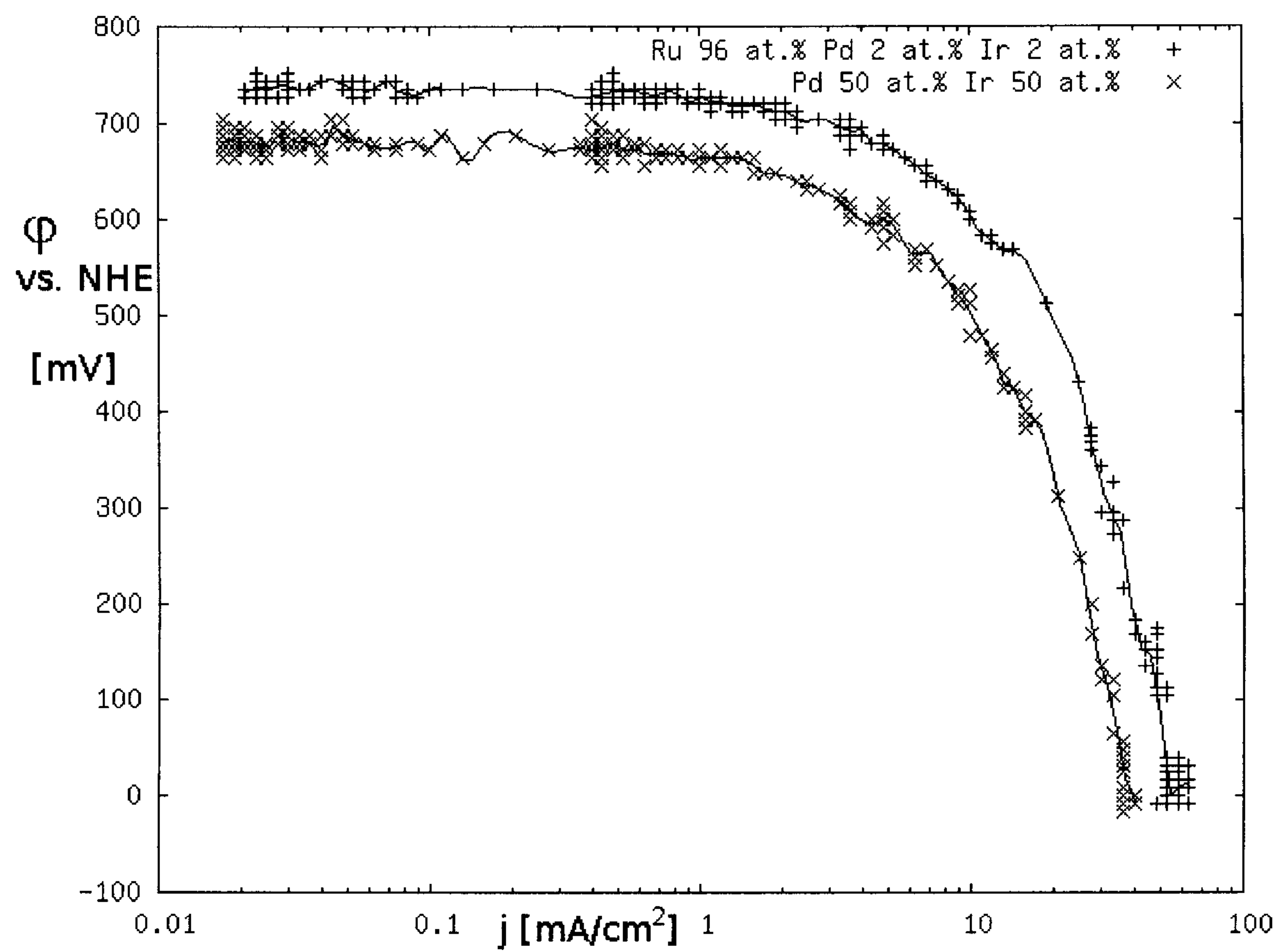


Fig. 7

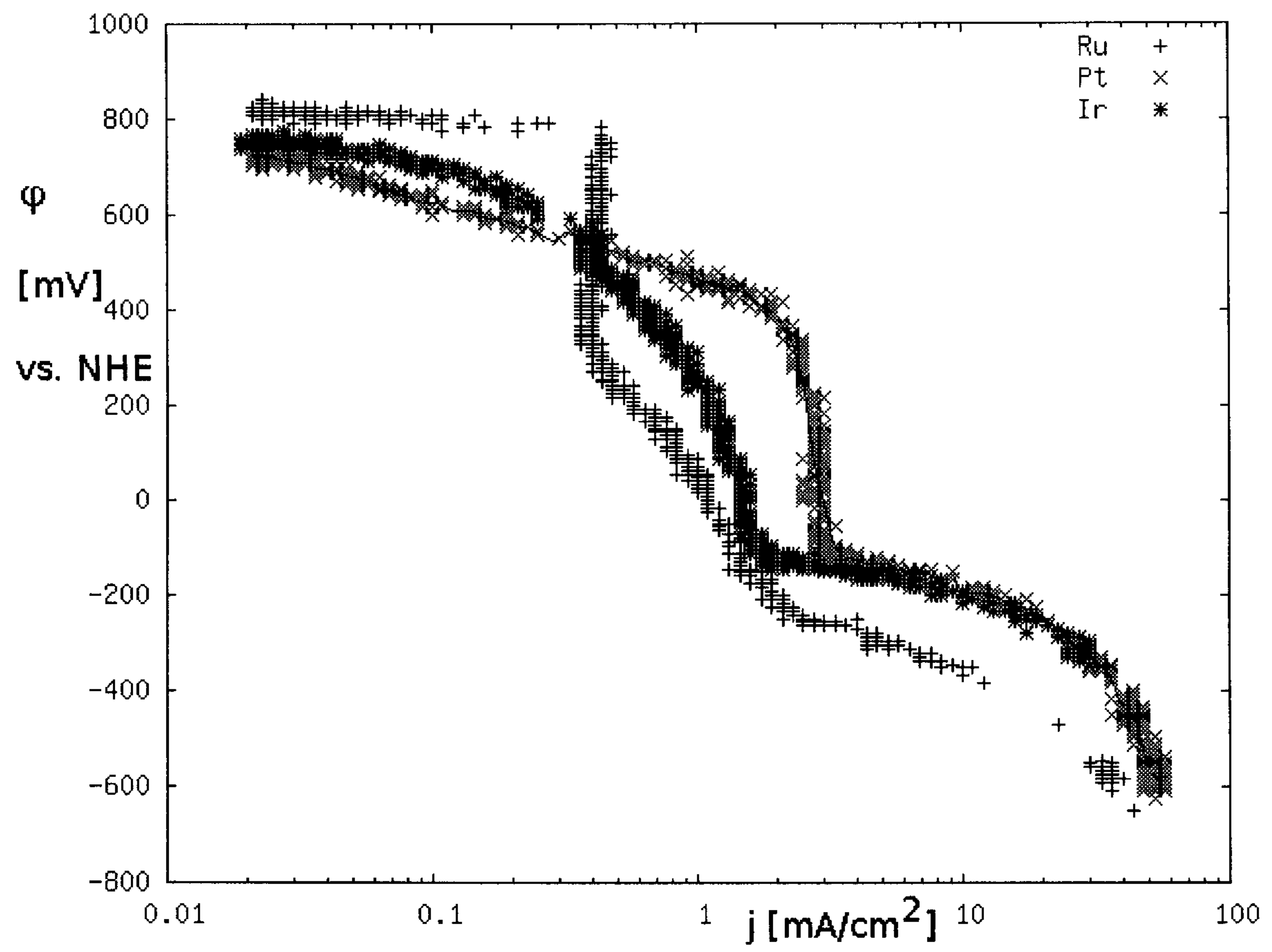


Fig. 8

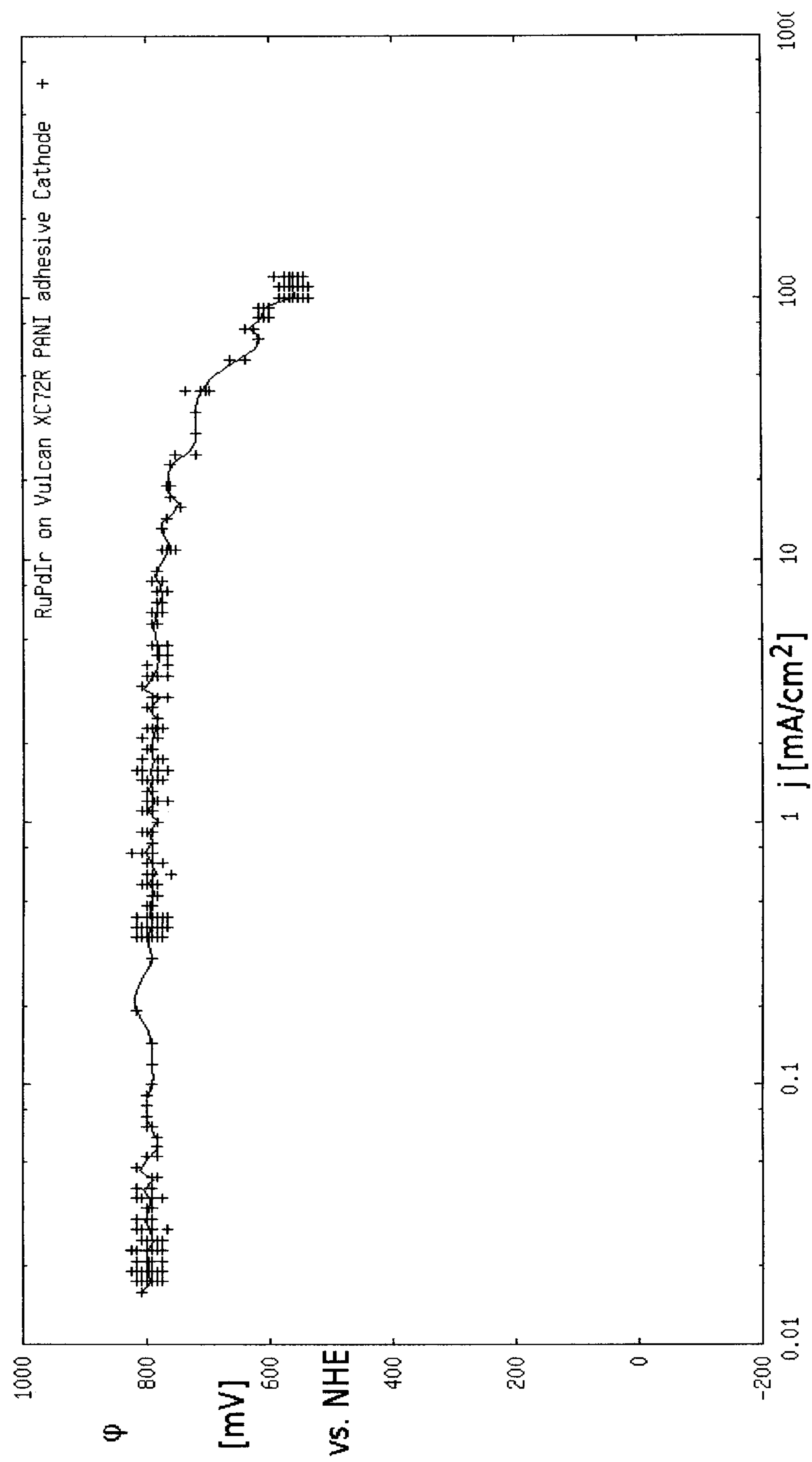


Fig. 9

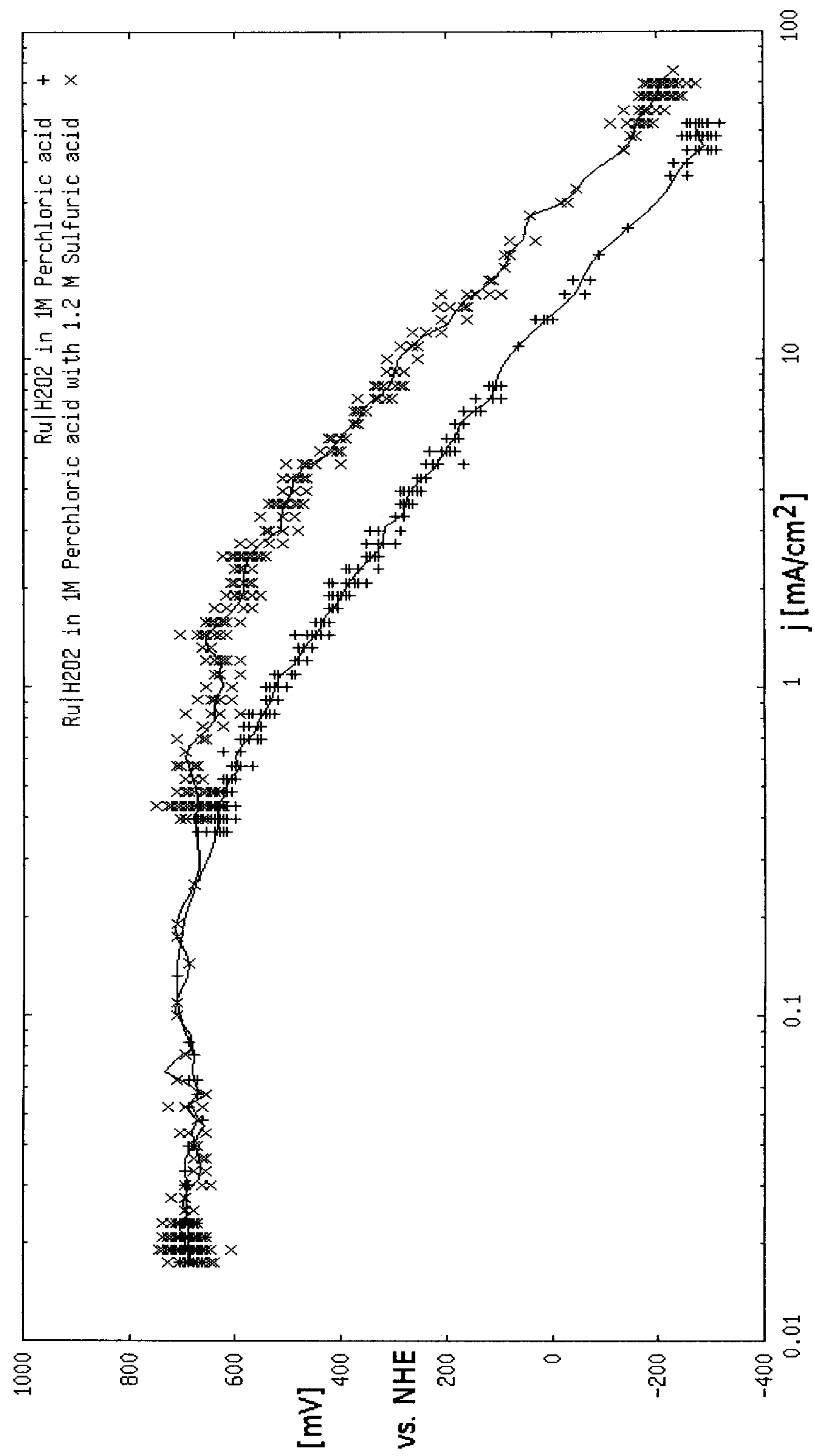


Fig. 10

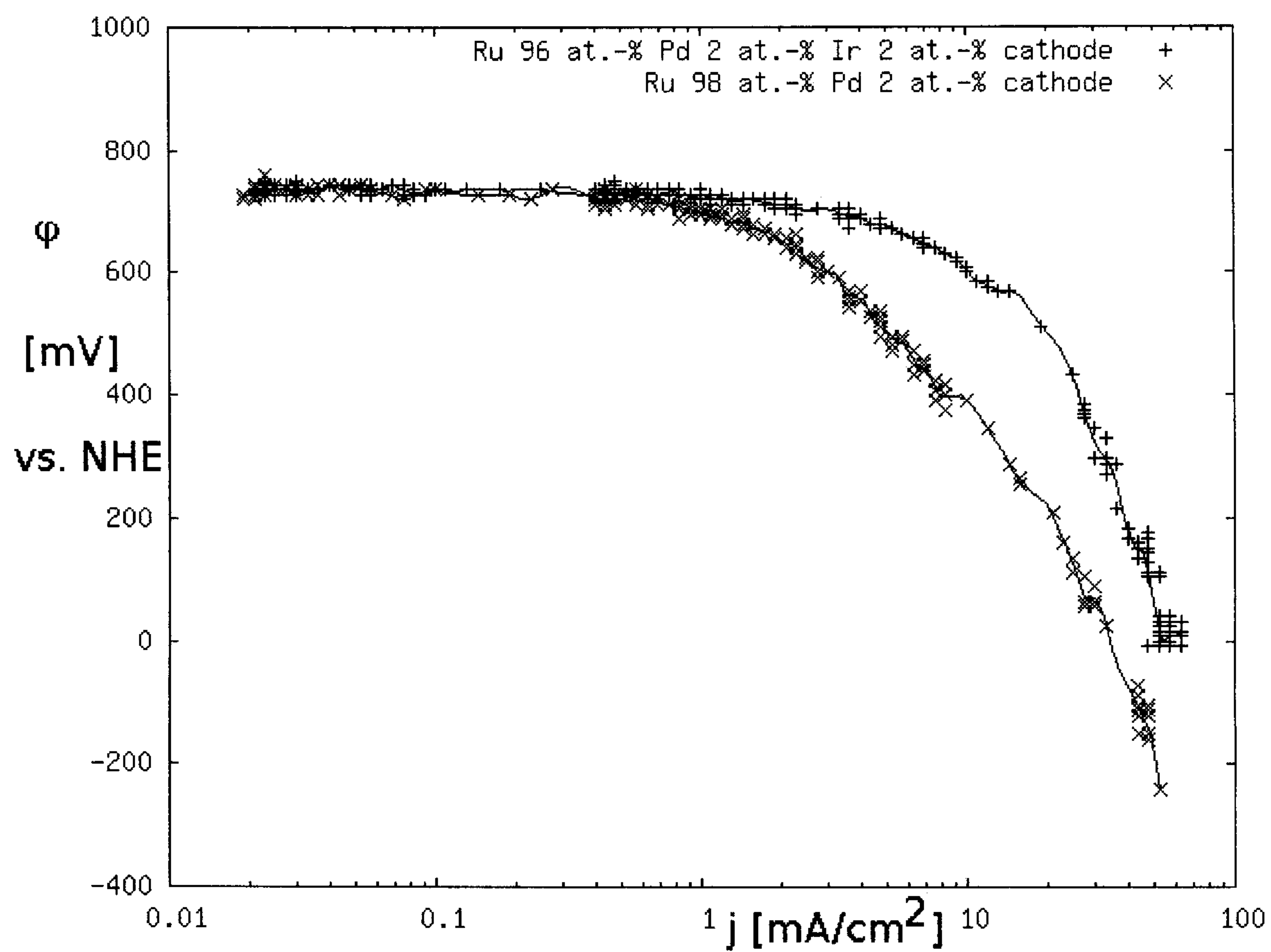


Fig. 11

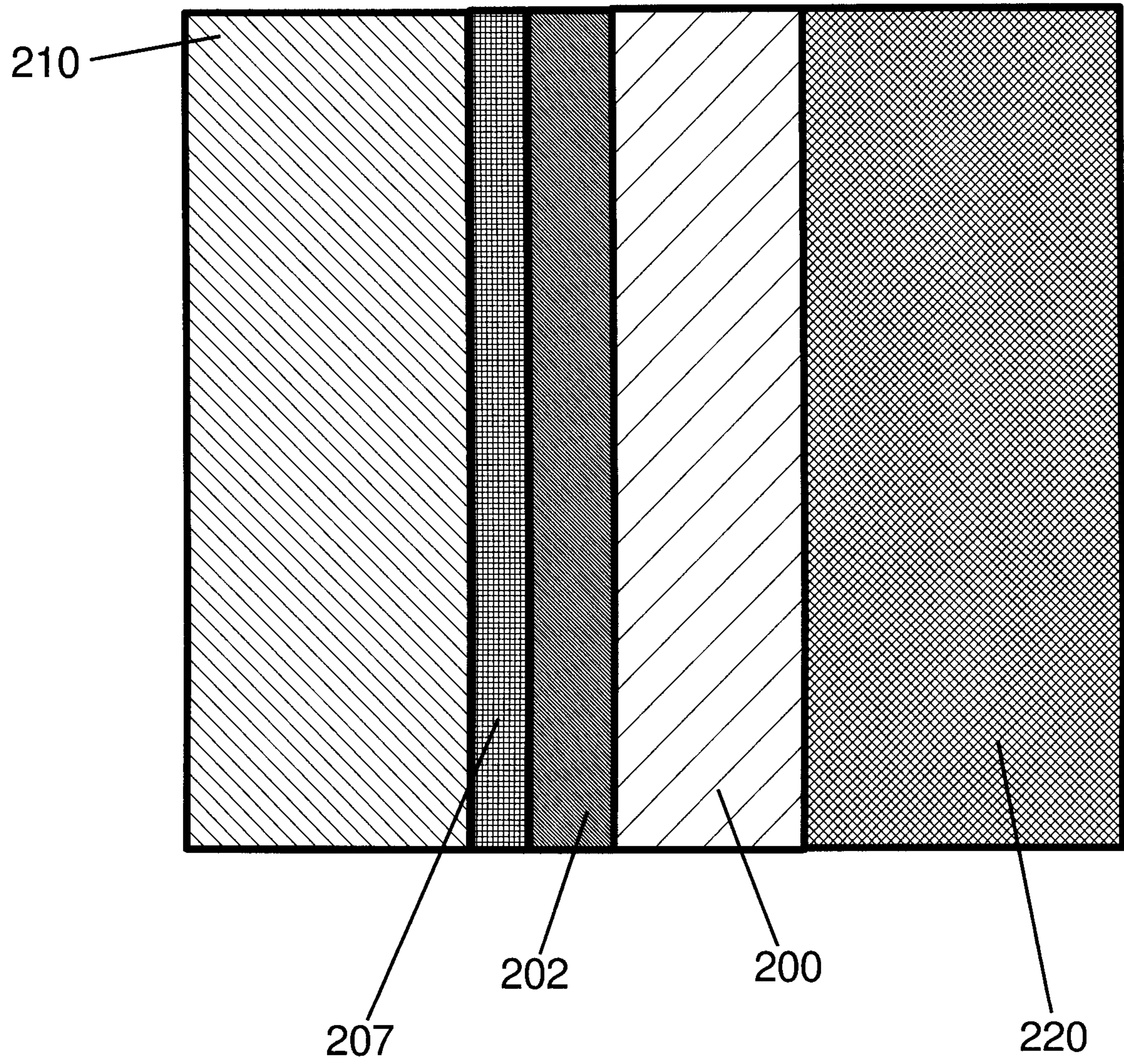


Fig. 12

