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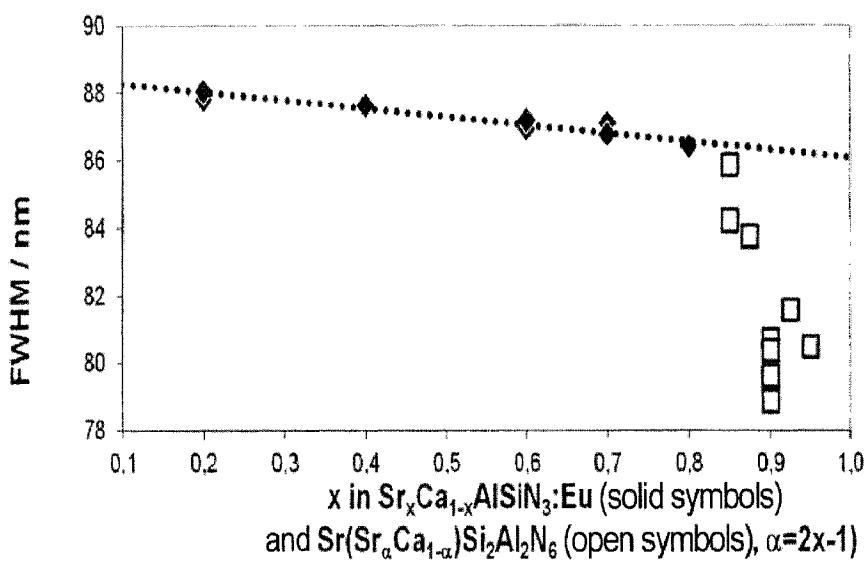
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(57) Abrégé/Abstract:

An embodiment of the invention relates to a luminescent material, comprising an inorganic substance which includes in its composition at least the element D, the element A1, the element AX, the element SX and the element NX (D representing one, two

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(57) Abrégé(suite)/Abstract(continued):

or more elements from the group consisting of Mn, Ce, Pr, Nd, Sm, Eu, Tb, Dy, Ho, Er, Tm and Yb, A1 representing one, two or more elements from the group of divalent metals which are not included in D, SX representing one, two or more elements from the group of tetravalent metals, AX representing one, two or more elements from the group of trivalent metals, and NX representing one, two or more elements from the group consisting of O, N, S, C, C1, F) and has the same crystal structure such as $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{A}_1\text{N}_6$.

Abstract

An embodiment of the invention relates to a luminescent material, comprising an inorganic substance which includes in its composition at least the element D, the element A1, the element AX, the element SX and the element NX (D representing one, two or more elements from the group consisting of Mn, Ce, Pr, Nd, Sm, Eu, Tb, Dy, Ho, Er, Tm and Yb, A1 representing one, two or more elements from the group of divalent metals which are not included in D, SX representing one, two or more elements from the group of tetravalent metals, AX representing one, two or more elements from the group of trivalent metals, and NX representing one, two or more elements from the group consisting of O, N, S, C, Cl, F) and has the same crystal structure such as $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{A1}_2\text{N}_6$.

Description

Phosphor, method for producing a phosphor and use of a phosphor

A phosphor is specified. Additionally specified are a process for producing such a phosphor and a use of such a phosphor. A phosphor suitable for use in semiconductor components such as LEDs or laser diodes is specified in EP 2 135 920 and EP 1 696 016.

One problem to be addressed is that of specifying a phosphor which has comparatively narrow-band spectral emission in the red spectral region.

This problem is solved inter alia by a phosphor, by a process and by a use having the features described herein.

It is generally a feature of a phosphor of the invention that the phosphor includes an inorganic substance which includes, in its composition, at least the element D, the element A1, the element AX, the element SX and the element NX (where D is one, two or more elements from the group of Mn, Ce, Pr, Nd, Sm, Eu, Tb, Dy, Ho, Er, Tm, alkali metals (Li, Na, K, Rb, Cs) and Yb, A1 is one, two or more elements from the group of the divalent metals not included in D, SX is one, two or more elements from the group of the tetravalent metals, AX is one, two or more elements from the group of the trivalent metals, and NX is one, two or more elements from the group of O, N, S, C, Cl, F) and has the same crystal structure as $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$.

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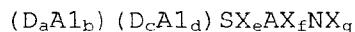
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The inventors have found that such a phosphor has a whole series of advantages over conventional phosphors, as described further down.

A phosphor "having the same crystal structure as $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$ " is defined hereinafter as a phosphor which, as well as the $\text{P}2_1$ space group, can additionally also be described in one of space groups 1 to 3 according to the International Tables Crystallography A, i.e. in the following space groups: $\text{P}1$, $\text{P}2$, $\text{P}\bar{1}$, and wherein the length of the chemical bonds between the elements Al-N and Si-N calculated from the lattice constants and atomic coordinates according to a Rietveld analysis is within a value of $\pm 15\%$ of the values described in figure 22.

In a further embodiment of the present invention, the space group is monoclinic $\text{P}2_1$.

In a further embodiment of the present invention, the inorganic substance can be described by the following general formula:



where $a+b \leq 1$ and $c+d \leq 1$ and where the parameters a, b, c, d, e, f and g fulfill the following conditions:
 $0 \leq a \leq 0.5$; $0 \leq c \leq 0.5$; $0 \leq b \leq 1$; $0 \leq d \leq 1$; $a+c > 0$;
 $b+d < 2$; $0.1 \leq e \leq 8$; $0.1 \leq f \leq 16$; $0.8(f + 4/3 e + 2/3 (b + d)) \leq g$; and $g \leq 1.2 (f + 4/3 e + 2/3 (b + d))$.

Preferably, the following conditions apply: $0 \leq a \leq 0.1$; $0 \leq c \leq 0.1$; $0 \leq b \leq 1$; $0 \leq d \leq 1$; $a+c > 0$; $b+d < 2$; $0.1 \leq e \leq 8$;

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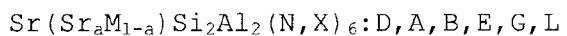
$0.1 \leq f \leq 16$; $0.8(f + 4/3 e + 2/3(b + d)) \leq g$; and $g \leq 1.2(f + 4/3 e + 2/3(b + d))$.

In a further embodiment, the phosphor has a general empirical formula $A_1(A_{1-a}M_{1-a})SX_2AX_2NX_6:D$. In this formula, A_1 is at least one divalent metallic element, for example Sr, M is another divalent metallic element, for example Ca, SX contains at least one tetravalent element, for example Si and/or C, AX contains at least one trivalent element, for example Al and/or La, and NX contains at least one element selected from the group of N, O, F, Cl.

In addition, the general elements A_1 , M , SX , AX and NX in this empirical formula may be defined as already described above, i.e. D as one, two or more elements from the group of Mn, Ce, Pr, Nd, Sm, Eu, Tb, Dy, Ho, Er, alkali metals (Li, Na, K, Rb, Cs), particularly Li, Tm and Yb, A_1 as one, two or more elements from the group of the divalent metals not included in D , SX as one, two or more elements from the group of the tetravalent metals, e.g. Si, C, Ge, Hf, Zr, Ti, AX as one, two or more elements from the group of the trivalent metals, e.g. Al, La, Ga, In, B, and NX as one, two or more elements from the group of O, N, S, C, Cl, F.

The parameter value a may be between 0.6 and 1.0, or between 0.8 and 1.0. In addition, it may be the case that $a < 1$.

The present invention further provides, in a further embodiment, a phosphor of the general formula:



where the phosphors of the invention are co-doped with the elements A, B, E, G and L and these co-dopants can occupy

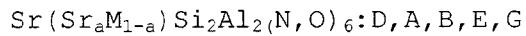
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either positions in the host lattice or interstitial lattice sites. The general element X represents elements such as O or halogen, for example, which especially also serve to balance charge carriers if occupied interstitial lattice sites are present in the crystal lattice or voids are present at lattice sites.

The metal M here is selected from Ca, Ba, Mg alone or in combination, A is selected from divalent metals which are different than M and than the further dopant D used, for example Cu, Zn or combinations thereof, B represents trivalent metals, especially transition or rare earth metals, e.g. La or Pr, and E represents monovalent metals, e.g. Li or other alkali metals such as Cs, Rb, K or Na. G represents tetravalent elements, for example C or Ge, or Hf, Zr, Ti. The element L here represents trivalent elements, for example B, Ga or In.

More particularly, this phosphor may have the following general formula:



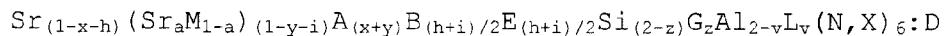
Hereinafter, rather than the general formula $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2(\text{N},\text{O})_6:\text{D},\text{A},\text{B},\text{E},\text{G}$, the formulae $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2(\text{N},\text{O})_6:\text{D}$ and $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2(\text{N},\text{O})_6:\text{Eu}$, $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2(\text{N})_6:\text{D}$ and $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2(\text{N})_6:\text{Eu}$ are used synonymously for reasons of simplicity.

The dopants may additionally occupy specific positions within the crystal lattice of the phosphors of the invention, for example lattice sites or interstitial lattice sites, and may also replace elements present in the phosphors, so as to

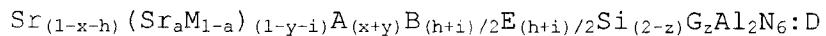
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result, in a further embodiment, in a phosphor of the following general formula:



More particularly, the general formula may be as follows:



where the metal M and the elements A, B and E are the elements just described above, and where $0 \leq x+y \leq 0.4$, preferably $0 \leq x+y \leq 0.3$, further preferably $0.04 \leq x+y \leq 0.3$, $x+y$ may especially be 0.2 or 0.04, and also $0 \leq h+i \leq 0.4$, preferably $0 \leq h+i \leq 0.3$, further preferably $0.04 \leq h+i \leq 0.3$, where it is also possible for no B to be present, such that $x = 0$ and $y = 0$. The parameters $h+i$ may especially be 0.2 or 0.04, where it is also possible for no B and E to be present, such that $h = 0$ and $i = 0$. In this case, the divalent metals A and/or a combination of equal molar proportions of the trivalent and monovalent metals B and E may replace Sr and/or Ca. The parameters $x+y$, $h+i$ and z may be selected independently of one another. In addition, it is also possible for each of x and y and h and i independently to be 0.

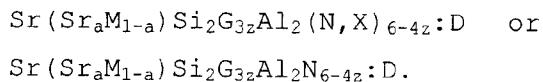
G represents tetravalent elements, for example C or Ge, which replace Si, where the parameter z is as follows: $0 \leq z \leq 1$, or $0 \leq z \leq 0.5$, or $0.02 \leq z \leq 0.3$, where z may especially be 0.02 or 0.4, or no tetravalent element may be present, such that $z = 0$. The parameter v for the element L may assume the following values: $0 \leq v \leq 1$, and also $0 \leq v \leq 0.5$.

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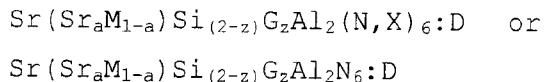
Replacement of Sr and M by A and/or replacement of Sr and M by a combination of B and E can lead to a change in the color locus in the CIE color space, to a change in the dominant wavelength, in the reflection capacity, in the luminous efficiency LER, in the thermal quenching characteristics, in the stability to radiation, in the hydrolysis sensitivity and/or in the FWHM of the phosphor of the invention, and hence open up further ways of adapting the phosphors of the invention for specific applications.

Replacement of Si by G can additionally lead to a significant shift in the wavelength of emission of the phosphor and can thus bring about an improvement in the color rendering index, particularly in the case of rendering of deep red colors. Co-doping with carbon, for example, thus increases the options for achieving particular color loci.

In addition, it is possible that the tetravalent element G, e.g. C, also partly replaces the nitrogen atoms in the phosphors of the invention, in which case G is present as G^{4-} , so as to result in the following general structural formula:



In a further embodiment of the phosphor, it is possible that each of $x+y$, $h+i$ and/or $z = 0$, in which case the following general formula is the result:

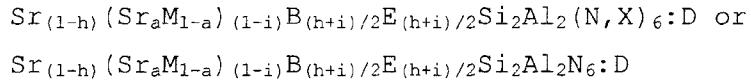


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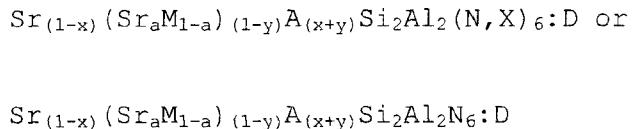
when $x+y = 0$ ($x = 0$ and $y = 0$) and additionally $h+i = 0$ ($h = 0$ and $i = 0$),

or the general formula:



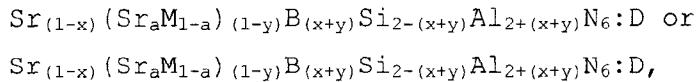
is the result when $z = 0$ and $x+y = 0$ ($x = 0$ and $y = 0$).

In addition, it is possible for exclusively divalent elements A to replace Sr and M i.e. for no B or E or G to be present, so as to result in the following general formulae:



In addition, M in the above formulae may preferably be Ca.

A further embodiment of a phosphor of the invention has the following general formula:



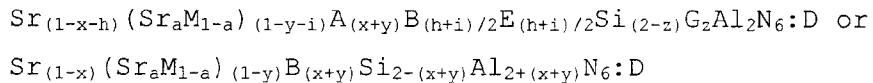
such that Sr and M, and Si are replaced by a combination of the trivalent metals B and Al, where the following applies here too: $0 \leq x+y \leq 0.4$, preferably $0.04 \leq x+y \leq 0.3$; $x+y$ may especially be 0.2.

All the above mentioned phosphors have strong absorption in the blue spectral region and emit red secondary radiation. In addition, these phosphors have the same crystal structure as

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$\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$, and so crystallize in the space groups P1, P2, $\bar{P}1$, especially in the monoclinic $P2_1$ space group.

A further embodiment of the present invention also provides phosphors of the general formula already described above:



where D is one, two or more elements from the group of Mn, Ce, Pr, Nd, Sm, Eu, Tb, Dy, Ho, Er, Tm, alkali metals, i.e. Li, Na, K, Rb, Cs, preferably Li and Yb. D is preferably selected from Eu, Ce, Li, Mn and combinations thereof.

Preferably, the activator D is selected from at least one element from Eu, Ce, Mn, especially Eu, Ce or Mn, or a mixture of Eu, Ce, Li. By using the latter activators, it is possible with particular efficiency to adjust the color locus of the phosphor in the CIE color space, its dominant wavelength λ_{dom} , the luminous efficiency LER, FWHM, and the diffuse reflectance at 450-470 nm.

Another particular possibility is co-doping of Eu-doped phosphors of the invention with alkali metals, i.e. Li, Na, K, Rb, Cs, preferably Li. Co-doping with the alkali metals may especially lead to a decrease in the spectral half-height width FWHM, and result in improved characteristics with regard to thermal quenching, and an improvement in the luminous efficiency.

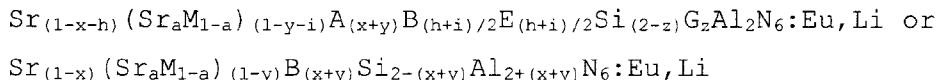
In a further embodiment of the present invention, the activator D is a combination of Eu and one or more alkali metals, preferably Li. This can lead to a further reduction

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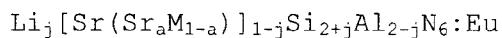
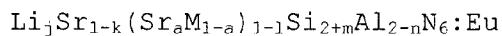
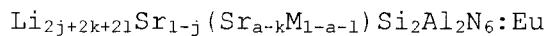
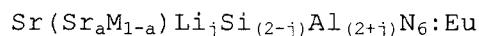
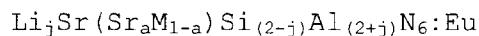
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in the FWHM of the radiation emitted, an improvement in the thermal quenching characteristics and the quantum efficiency.

A general formula of this phosphor having a combination of Eu and Li can be described as follows:



In this case, it is possible that the lithium metal ions occupy positions in the host lattice and/or are present at intermediate lattice sites. Balancing of charge carriers is possible by adjusting the Si:Al ratio and/or by partly replacing N with O and/or halogens such as F. Also possible are randomly distributed voids within the cation lattice sites. For this reason, the following general formulae are also suitable for description of phosphors of the invention having Eu and Li as activators, without showing the additional elements A, B, E and G for reasons of clarity, but these may fundamentally be present:



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The following applies to the parameter j : $0 \leq j \leq 0.2$, preferably $0 \leq j \leq 0.15$, further preferably $0 \leq j \leq 0.05$. In a further embodiment, the phosphor has a general empirical formula $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{D}$. In this formula, M is Ca and/or Ba. In addition, M may also be selected from the group of Ca, Ba, Zn, Mg and/or Li, alone or in combination. In these cases and the above-described general formulae, the value of a may be 0.6 to 1.0, preferably 0.8 to 1.0 (excluding the boundary values). More particularly, a is chosen at 0.7 to 0.99, further preferably at 0.85 to 0.99, including the boundary values.

The activator D may, in a further embodiment of the invention, be present in molar percentages between 0.1% and 20 mol%, or 0.1% and 10%, or 1 mol% - 10 mol%, or 0.5% to 5%, 2-5 mol%, or 0.8% to 3%. Here and hereinafter, percentage figures for the activator, especially Eu, are understood as molar percentages based on the molar proportions of the alkaline earth metals in the particular phosphor.

The activator D may preferably be selected from the metals Eu, Ce, Mn and Li and combinations thereof. In addition, the activator D may be Eu, Mn or Ce, and combinations of Eu, Ce and Li.

With the rising concentration of europium dopant, the dominant wavelength of the emission of the phosphors of the invention moves to higher wavelengths from the orange to the red color region (see figure 58a), with a rise in the relative intensity of photoluminescence of 0.1 to about 4 mol% and then a drop again as the activator concentrations of europium continue to rise (see figure 58b). Based on the relative intensity of the photoluminescence, a concentration

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range of 1-10 mol% Eu, or 2-5 mol%, is preferred. Behavior substantially analogous to the relative intensity of photoluminescence is also displayed by the relative photoluminescence intensity assessed by eye, which likewise increases with rising activator concentrations of europium and drops again from about 4 mol% to about 20 mol% (see figure 58c). Based on the photoluminescence intensity assessed by eye, activator concentrations of 0.4-10 mol% of Eu, or 1-5 mol% of europium, are preferred.

In at least one embodiment, the phosphor is set up to emit red or orange light. Red or orange light means that the phosphor emits radiation having a dominant wavelength of at least 560 nm, preferably between 585 nm and 640 nm inclusive, especially between 590 nm and 615 nm inclusive.

The dominant wavelength is especially that wavelength which is found to be the point of intersection of the spectral color line of the CIE standard chromaticity diagram with a straight line, this straight line proceeding from the white point in the CIE standard chromaticity diagram and running through the actual color locus of the radiation. In general, the dominant wavelength differs from a wavelength of maximum intensity. More particularly, the dominant wavelength in the red spectral region is at smaller wavelengths than the wavelength of maximum intensity.

In at least one embodiment, the phosphor has a general empirical formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{D}$. In this formula, D is at least one activating element. Frequently, D is formed by element Eu and/or else Ce. Other or additional activating elements or dopants may be selected from the group of Mn, Ce,

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Pr, Nd, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, and Lu, each alone or in combination. It is possible that the phosphor includes further elements, for instance in the form of impurities, in which case these impurities taken together should preferably have a maximum proportion by weight in the phosphor of not more than 0.1 permille or 100 ppm or 10 ppm, parts per million.

In at least one embodiment, the phosphor is set up to emit red light and preferably to be excited by blue light. Phosphors which can be excited in the ultraviolet spectral range into the blue-green spectral range and which emit red light are of great significance for the production of white light-emitting diodes. Specifically in the case of light-emitting diodes having low color temperatures, called warm white light-emitting diodes, and/or in the case of light-emitting diodes having a high color rendering index, phosphors of this kind are required. Phosphors of this kind are also required in a multitude of other applications, for instance for display backlighting, what are called color-on-demand applications or else for orange and red full conversion light-emitting diodes. Use in combination with an organic light-emitting diode, OLED for short, is likewise possible. The phosphor described here is usable for such applications, and likewise for laser applications such as the LARP method.

In at least one embodiment, the phosphor in a powder diffractogram on irradiation with monochromatic $\text{Cu-K}_{\alpha 1}$ radiation has a reflection at an angle 2 theta between 36.7° and 37.0° , according to the composition of the phosphor. The exact position of this reflection depends on the general empirical formula of the phosphor. An intensity of this

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reflection, especially based on a main reflection, is preferably at least 0.3% or 0.5% and/or at most 10% or 8% or 5% or 4%.

In at least one embodiment, the dominant wavelength of the phosphor is at least 596 nm or 598 nm. Alternatively or additionally, the dominant wavelength is at most 610 nm, 606 nm or 604 nm. The wavelength of maximum intensity is, for example, at least 605 nm or 610 nm and/or at most 630 nm or 625 nm.

In at least one embodiment, the phosphor has a spectral half-height width at half the maximum height, FWHM or full-width at half maximum for short, of at least 70 nm or 75 nm or 78 nm. The maximum of this spectral range is preferably 90 nm or 87 nm or 84 nm or 82 nm.

In at least one embodiment, the phosphor has an absorption maximum in the blue spectral region, especially a relative absorption maximum. The blue spectral region especially refers to wavelengths of at least 400 nm and/or of at most 480 nm. For example, the absorption maximum is at at least 410 nm or 420 nm and/or at at most 450 nm or 440 nm.

The abovementioned values relating to the spectral properties of the phosphor especially apply at room temperature, i.e. at about 300 K.

Additionally specified is a process for producing such a phosphor. Features of the phosphor are therefore also disclosed for the process, and vice versa.

In at least one embodiment, the process has at least the following steps, preferably in the sequence specified:

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- A) providing reactants in the solid state for Sr, Al, Si and Eu and optionally for Ca,
- B) mixing the reactants,
- C) heating the reactants under an inert gas atmosphere, especially nitrogen atmosphere, or under a forming gas atmosphere, to at least 1500°C and forming a calcined cake, and
- D) comminuting the calcined cake to give the phosphor.

In at least one embodiment of the process, at least step C) or else all steps are effected at about atmospheric pressure. More particularly, the process in that case is not effected under high pressure conditions. Preferably, the atmospheric pressure and/or a total pressure is between 0.9 bar and 1.5 bar or 0.95 bar and 1.05 bar inclusive.

Reactants and sources used for strontium, aluminum and/or calcium may be the respective pure metals or else metal alloys with the appropriate metals. Reactants used may likewise be silicides, hydrides, nitrides, oxynitrides, halides and/or oxides of these metals. In addition, it is possible to use mixtures of these compounds.

Reactants or sources used for silicon for the production of the phosphor may be a silicon-metal compound, a silicon nitride, an alkaline earth metal silicide, silicon diimide, or a mixture of these compounds. Preference is given to using silicon nitrides and/or silicon metals.

Reactants or sources used for Eu may be metallic europium, a europium alloy, a europium oxide, a europium nitride, europium hydride or a europium halide. It is likewise

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possible to use mixtures of these compounds. Preference is given to using europium oxide as reactant for europium.

Reactants used for the further tetravalent elements G, e.g. C, the trivalent elements B, e.g. La, the monovalent elements E, e.g. Li, and the divalent elements A other than D and M, e.g. Cu or Zn, may, for example, be the corresponding elements, silicides, hydrides, nitrides, oxynitrides, carbonates, hydroxides, halides and/or oxides of these elements and compounds derived therefrom, for example hydrates. For example, it is possible to use Mn_2O_3 , CuO , Zn_3N_2 , La_2O_3 , $Li_2B_4O_7$ and graphite.

In at least one embodiment, a melting agent and/or a flux is used for the improvement of crystallinity and/or to promote crystal growth. For this purpose, preferably chlorides, fluorides, halides and/or boron compounds of the alkaline earth metals used are employed. Combinations of two or more melting agents or fluxes may also be used. Melting agents and/or fluxes used are especially, for example, at least one of the following substances: LiF , $LiCl$, NaF , $NaCl$, $SrCl_2$, SrF_2 , $CaCl_2$, CaF_2 , $BaCl_2$, BaF_2 , NH_4Cl , NH_4F , KF , KCl , MgF_2 , $MgCl_2$, AlF_3 , H_3BO_3 , B_2O_3 , $Li_2B_4O_7$, $NaBO_2$, $Na_2B_4O_7$, $LiBF_4$. Also suitable are NH_4HF_2 , $NaBF_4$, KBF_4 , EuF_3 and compounds derived therefrom, for example hydrates.

In at least one embodiment, the starting substances, especially for Sr, Ca, Al and/or Si and also Eu, and optionally also for the further tetravalent elements G, e.g. C, the trivalent elements B, e.g. La, the monovalent elements E, e.g. Li, and the divalent elements A other than D and M, e.g. Cu or Zn, are weighed out according to the general empirical formula of the phosphor. It is possible that the

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alkaline earth metal components Sr, Ca are also weighed out with an excess, in order to compensate for any evaporation losses that occur during the synthesis. In addition, it is also possible to employ Ba as alkaline earth metal component.

In at least one embodiment, step D) is followed by a step E).

In step E), the phosphor is calcined further, which can also be referred to as tempering. The calcination is especially effected at a temperature of at least 1500°C and preferably under a nitrogen atmosphere or forming gas atmosphere.

Forming gas refers to a mixture of N₂ and H₂. The temperature of at least 1500°C in steps C) and/or E) is preferably employed for at least four hours or six hours. For example, in each of steps C) and E), a temperature of 1650°C ± 50°C is employed.

In an alternative embodiment of a process of the invention for preparation of such a phosphor, rather than step E), it is also possible to repeat steps C) and D).

In at least one embodiment, the reactants are mixed in a ball mill or in a tumbling mixer. In the mixing operation, it may be advantageous to choose the conditions such that a large amount of energy is introduced into the mixture, which results in grinding of the reactants. The resultant increase in homogeneity and reactivity of the mixture can have a positive influence on the properties of the resulting phosphor.

By controlled alteration of the bulk density or by modification of the agglomeration of the reactant mixture, it is possible to reduce the formation of secondary phases. In addition, a particle size distribution, a particle morphology

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and/or a yield of the resulting phosphor can be affected. Techniques of particular suitability for the purpose are screening and pelletizing operations, including use of suitable additives.

In at least one embodiment, a tempering is effected, especially in a crucible made from tungsten, molybdenum or boron nitride. The tempering is preferably effected in a gas-tight oven in a nitrogen atmosphere or in a nitrogen/hydrogen atmosphere. The atmosphere may be flowing or stationary. It is additionally possible for carbon in finely divided form to be present in the oven space. Also possible are multiple tempering steps of the phosphor, in order to improve or to control the crystallinity or particle size distribution. Further advantages may be a lower defect density, associated with improved optical properties of the phosphor and/or a higher stability of the phosphor. Between the heat treatments, the phosphor may be treated in a wide variety of different ways, or it is possible to add substances such as melting agents to the phosphor.

For grinding of the phosphor, it is possible, for instance, to use a mill, a fluidized bed mill or a ball mill. In the grinding operation, it is preferable to ensure that the proportion of splintered grains produced is kept to a minimum, since these can worsen the optical properties of the phosphor.

The phosphor can additionally be washed. For this purpose, the phosphor can be washed in water or in aqueous acids such as hydrochloric acid, nitric acid, hydrofluoric acid, sulfuric acid, organic acids or a mixture of these. The phosphor may alternatively or additionally be washed in an

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alkali such as sodium hydroxide solution, potassium hydroxide solution, an aqueous ammonia solution or mixtures thereof. Alternatively or additionally, washing in organic solvents such as acetone, propanol and/or phenol is possible. The washing preferably follows after the grinding.

In at least one embodiment, for instance, the empering, further calcining, grinding, sieving and/or washing result in removal of secondary phases, glass phases or other contamination and hence an improvement in the optical properties of the phosphor. It is also possible by this treatment to selectively remove or dissolve small phosphor particles and to influence the particle size distribution for the application. In addition, such a treatment can alter a surface of the phosphor particles in a controlled manner, for example the removal of particular constituents from the particle surface. This treatment can, also in conjunction with a downstream treatment, lead to improved stability of the phosphor. More particularly, the application of a protective layer is possible, as is basically known per se.

Additionally specified is the use of such a phosphor. Features relating to use are therefore also disclosed for the process and the phosphor, and vice versa.

In at least one embodiment, the phosphor is used in a light-emitting diode as radiation source as the first phosphor in a lighting device. The light-emitting diode comprises at least one semiconductor chip that emits in the blue and/or UV spectral region in operation. The phosphor is arranged downstream of the semiconductor chip along a beam path.

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The blue and/or UV light produced by the semiconductor chip is partly or fully absorbed by the phosphor and converted to radiation of greater wavelength, especially to red (e.g. D = Eu) or orange (e.g. D = Ce) light. It is possible that at least one further second phosphor which has an emission different than the first phosphor and is especially suitable for generation of green and/or yellow light is present, especially also phosphors having the same structure as the first phosphor. In addition, the light-emitting diode preferably emits mixed radiation including blue light from the semiconductor chip and converted radiation from the phosphor, and also green and/or yellow light from the further phosphor. The primary radiation used may also be UV radiation rather than blue light.

As well as the first phosphor and optionally the second phosphor, it is also possible for further non-luminous particles, for example scattering particles and diffusers, to be present in the beam path of the radiation source.

In the remarks which follow, the composition of the novel phosphor of the invention having the empirical formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ is described. This corresponds to the nominal composition of the samples according to the composition weighed out. However, the Si:Al ratio actually present may also differ from 2:2, which can be expressed by a parameter d. A possible value of d is, for example, between 0 and 0.1, including the boundary values. No distinction of Si and Al by x-ray methods is possible. It is likewise possible that the finished phosphor contains other, further elements which are introduced, for example, via impurities or fluxes in the starting mixture or during the synthesis, especially (but not exclusively) boron and/or carbon and/or oxygen

and/or halogens, for example fluorine or chlorine. By virtue of possible evaporations of individual components, there may also be statistical underoccupation of individual positions. This effect too can be detected only with difficulty, if at all, by x-ray analysis.

Therefore, in all embodiments, an empirical formula of the $\text{Sr}_{1-e}(\text{Sr}_a\text{Ca}_{1-a})_{1-g}(\text{Si},\text{C})_{2+d}(\text{Al},\text{B})_{2-d}(\text{N},\text{O},\text{F},\text{Cl},\text{C})_6:\text{Eu}$ type is accordingly a possible description of the phosphor actually present.

For reasons of clarity, hereinafter, however, all embodiments will refer simply to "Sr(Sr_aCa_{1-a})Si₂Al₂N₆:Eu"; in that case, the particular empirical formula specified corresponds to the composition weighed out.

In a further embodiment, there is provided a phosphor for emission of red light, having the general empirical formula $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ wherein $0.8 < x \leq 1$, wherein between 0.1% and 5% inclusive of the Sr, Ca and/or Sr/Ca lattice sites have been replaced by Eu, wherein a full-width at half maximum is between at least 70 nm and at most 90 nm, and wherein, in an x-ray structure analysis, the phosphor in orthorhombic description exhibits a reflection (R) having the Miller indices **121**.

In a further embodiment, there is provided a process for producing a phosphor as described herein, having the steps of: A) providing reactants in the solid state for Sr, Al, Si and Eu and optionally for Ca, B) mixing the reactants, C) heating the reactants under a nitrogen atmosphere or under a forming gas atmosphere to at least 1500°C and forming a calcined cake, and D) comminuting the calcined cake to give the phosphor.

In a further embodiment, there is provided a light-emitting diode comprising: at least one semiconductor chip configured to emit blue light; a phosphor arranged downstream of the semiconductor chip along a beam path wherein the phosphor emits red light during operation and has the general empirical formula: $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ wherein x is $0.8 < x \leq 1$, wherein between 0.1% and 5% inclusive of the Sr, Ca and/or Sr/Ca lattice sites have been replaced by Eu, wherein a full-width at half maximum is between at least 70 nm and at most 90 nm, and wherein the phosphor in orthorhombic description exhibits a reflection (R) having the Miller indices **121**.

A phosphor described here is elucidated in detail hereinafter by embodiments with reference to drawings. Identical reference numerals indicate identical elements in the individual figures. However, the drawings are not to scale; instead, individual elements may be shown in excessively large size for better understanding.

The figures show:

figure 1 relationship between x and FWHM;

figures 2 to 34 schematic diagrams of properties of phosphors or phosphor mixtures described here on excitation with blue light and data from an x-ray structure analysis of phosphors or phosphor mixtures described here compared to the prior art;

figures 35a to 36b emission spectra and reflectance spectra of phosphors described here;

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figure 37 a semiconductor component which serves as light source (LED) for white light;

figure 38 a lighting unit comprising phosphors according to the present invention;

figures 39a and 39b starting weights m in g for various reactants for production of conventional and inventive phosphors;

figures 40a to 44b and 45a to 45d show the luminous flux and the radiant power and the composition of the phosphors of various LEDs according to comparative examples and inventive examples of the present invention and the corresponding emission spectra thereof and data derived from the spectra for the full conversion of the primary radiation to red secondary radiation;

figures 46a to 46e and 47a to 47i show various embodiments of lighting devices of the invention that are suitable for streetlighting applications and the optical properties thereof;

figures 48a to 48j show embodiments of lighting devices of the invention for backlighting applications and the optical properties thereof;

figures 49a to 49g and 50a to 50e show experimental data for various comparative and inventive examples of lighting devices for flash applications;

figures 51a to 51h, and 52a to 52h and 53a to 53d show experimental data for lighting devices according to various comparative and inventive examples for warm

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white general-purpose lighting applications with high CRI;

figures 54 to 57 show the light yield and the composition of the phosphors for various LEDs in LARP applications according to comparative examples and inventive examples of the present invention and the corresponding emission spectra thereof;

figures 58a to 58c show the effects of different Eu dopant concentrations on the dominant wavelength, the relative intensity of photoluminescence and the relative photoluminescence intensity assessed by eye;

figures 59a to 59g show the composition of phosphors of the invention which have been co-doped with Cu, Zn, La and Li, and their spectra and x-ray diffractograms;

figures 60a and 60b show the nominal composition and the spectra of phosphors of the invention which have been co-doped with carbon;

figures 61a to 61d show the nominal composition and the spectra of phosphors of the invention which have been doped with various activators, including europium, cerium, lithium and manganese;

figures 62a to 62e show various properties of phosphors of the invention and the x-ray diffractograms thereof, these having been co-doped not only with europium but also with lithium;

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figures 63 to 73 show various embodiments of lighting devices
30 comprising the phosphors of the invention in
cross section.

One embodiment of a phosphor described here can be produced
as follows:

Reactants used for the synthesis of the phosphor of the
general empirical formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ are the
binary nitrides of the constituent elements, i.e. Ca_3N_2 ,
 Sr_3N_2 , AlN and Si_3N_4 . Since these are highly oxidation-
sensitive and hydrolysis-sensitive substances, what is called
a glovebox is employed, under an N_2 atmosphere with $\text{O}_2 < 1$
ppm and $\text{H}_2\text{O} < 1$ ppm. In addition, doping with Eu^{2+} is
accomplished using Eu_2O_3 . The reactants are weighed out such
that the following atomic ratio is effectively present, in a
simplifying representation:

$\text{Sr} : \text{Ca} : \text{Si} : \text{Al} : \text{Eu} = (1+a) : (1-a) : 2 : 2 : y$, where y
corresponds to the degree of doping, i.e. the proportion of
divalent lattice sites which are substituted by Eu. In
addition, various fluxes are added; see the above
explanation. A reactant mixture is scaled up, for example, to
a total starting weight of 50-100 g with retention of the
atomic ratios described above. It is also possible to use
other total starting weights.

The reactant mixture is introduced, for example, into a PET
mixing vessel together with ZrO_2 balls and mixed on a roller
table in a glovebox for 6 h. Subsequently, the balls are
removed from the mixture and the powder is transferred into a
closed molybdenum crucible. This crucible is placed into an
outer tungsten crucible, a semicircular open tungsten tube,

and transferred into a tube furnace. There is a flow of 3 l/min of forming gas with 92.5% N₂ and 7.5% H₂ through the tubular furnace during the run time. In the tubular furnace, the mixture is heated at a rate of 250 K/h to 1650°C, kept at this temperature for 4 h and then cooled down at 250 K/h to 50°C. The calcined cake formed is taken out after the furnace has cooled, comminuted with a mortar mill and sieved through a sieve having a mesh size of 31 µm. The sieve fraction < 31 µm is the phosphor used.

The sieving may optionally be followed by a further calcining, tempering and/or washing operation and/or a coating operation.

Illustrative starting weights m in g and resulting color loci CIE x, CIE y, also referred to as chromaticity coordinates, of the emission spectrum of the particular phosphor in the CIE standard chromaticity diagram on excitation with blue light at 460 nm and on complete absorption of the blue light are listed in tabular form in figures 39a and 39b. Starting weights with x ≤ 0.8 refer here to conventional phosphors, whereas starting weights with x > 0.8 (corresponding to a > 0.6) indicate phosphors of the invention.

Figure 1 shows the relationship between x and FWHM. Figure 1 shows a dependence of a spectral half-height width FWHM of the emission on the Sr content for embodiments of the novel phosphor Sr(Sr_aCa_{1-a})Si₂Al₂N₆:Eu of the invention (open symbols) compared to known phosphors composed of the Sr_xCa_{1-x}AlSiN₃:Eu system (solid symbols). The Sr content corresponds to the parameter a for Sr(Sr_aCa_{1-a})Si₂Al₂N₆:Eu or x for Sr_xCa_{1-x}AlSiN₃:Eu, where a = 2x-1. The parameter a is indicated by alpha in the figure.

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For the phosphor according to the prior art, a very small change in the FWHM with rising x is observed (dotted line). The novel phosphor of the invention, in contrast, at an Sr content of $a > 0.6$, which would correspond in formal terms to an $x > 0.8$ in the case of the known phosphor, exhibits a very significant change in the half-height width FWHM with rising a . Moreover, the half-height width of the novel phosphor is significantly lower than in the case of the phosphor according to the prior art. The parameter a can therefore readily be chosen between 0.6 and 1.0, or between 0.8 and 1.0; boundary values are preferably excluded. Very good properties are exhibited especially by phosphors having a value of a between 0.64 and 0.96, or between 0.82 and 0.98, including the boundary values. Particular preference is given to a range for a between 0.68 and 0.92, or between 0.84 and 0.96, including the boundary values.

In the case of use of constituents other than Sr, Ca, the value of a may also be much lower.

Figure 2 shows the ratio between l_{dom} (dominant wavelength) and the FWHM for various phosphors.

What is shown here is a dependence of a spectral half-height width FWHM of the emission on the Sr content for embodiments of the novel phosphor $Sr(Sr_aCa_{1-a})Si_2Al_2N_6:Eu$ of the invention (open symbols) compared to known phosphors composed of the $Sr_xCa_{1-x}AlSiN_3:Eu$ system (solid symbols). The Sr content corresponds to the parameter a for $Sr(Sr_aCa_{1-a})Si_2Al_2N_6:Eu$ or x for $Sr_xCa_{1-x}AlSiN_3:Eu$, where $a = 2x-1$. In addition, a dominant wavelength l_{dom} of the spectrum emitted by the phosphor and the Eu content are specified.

Surprisingly, a phosphor $Sr(Sr_aCa_{1-a})Si_2Al_2N_6:Eu$ of the invention with $a = 0.8$ (which would correspond to $x = 0.9$),

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with a comparable dominant wavelength λ_{dom} , shows a much smaller half-height width FWHM compared to conventional phosphors of the $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ type. The much smaller half-height width FWHM is virtually independent of the Eu content of the sample used.

Figure 3 gives a comparison of spectral data.

Phosphors of the novel type feature a small half-height width FWHM of the emission compared to the previously known phosphor (within the range of 79 to 81 compared to 86 to 88 for the prior art) and a very high luminous efficiency LER (110 to 111% compared to 100 to 101% for the prior art) combined with simultaneously high internal quantum efficiency QI and external quantum efficiency QE (about 113% compared to 100% to 110% for the prior art); see the table in figure 3. In addition, a relative brightness B is specified (about 125 to 126% compared to about 100 to 111% for the prior art). To calculate the external quantum efficiency QE, the mean re-emission within the range from 450 nm to 470 nm was employed; measurement was effected in pressed powder tablets at an excitation wavelength of 460 nm. Also specified are the x and y components of the color locus.

Figure 4 compares the LED efficiency of various systems. The relative conversion efficiencies of various warm white light-emitting diodes, LEDs for short, are shown. In each case, a mixture of two phosphors that emit green and red light was used, with the green light-emitting phosphor G remaining the same (such phosphors, especially garnets doped with Ce, are known per se) and the red-emitting phosphor R being varied. Stated on the abscissa axis for four different phosphors is the type of red-emitting phosphor R. The ordinate axis gives the relative efficiency E. The phosphors were excited with a

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blue-emitting semiconductor chip having a dominant wavelength of 446 nm.

All the phosphor mixtures were adjusted such that, in the CIE standard chromaticity diagram, a color locus close to Planck with a correlated color temperature CCT of about 2700 K is attained. The color rendering index CRI of all the LEDs measured is 80 ± 1 . All the red phosphors R used show a comparable dominant wavelength of about $600.5 \text{ nm} \pm 1 \text{ nm}$.

Further details of the phosphor mixtures as shown in figure 4 can be found in the table in figure 5. Additionally stated are the relative efficiency E (rel. eff.), a phosphor concentration c (phosphor concentration) and a ratio V of the green phosphor G and the red phosphor R (ratio green/red).

Figure 6 shows a comparison of conversion efficiency and color rendering index for various warm white light-producing LEDs. In each case, a mixture of two phosphors was used, with the green phosphor G being kept constant and the red phosphor R being varied, analogously to the table in figure 5. All the phosphor mixtures were adjusted such that a color locus close to Planck having a correlated color temperature CCT of about 2700 K is achieved. The relative conversion efficiency E (left-hand ordinate), the relative size of which is illustrated by the columns in figure 6, of a warm white light-producing LED with the novel phosphor having a total of 90% Sr at the alkaline earth metal site (shown on the right) shows a much higher efficiency (about 6% compared to a 258-nitride) and simultaneously improved color rendering CRI (right-hand ordinate, the color rendering index is symbolized as black rhombuses) compared to LEDs having a previously known red phosphor with only 80% Sr (1113-calsin type) or an

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even lower Sr content (258-nitridosilicate type) at the alkaline earth metal site.

Further data relating to the LED measurements from figure 6 can be found in the table in figure 7, analogously to the table in figure 5. The efficiency E (rel. eff.) of a warm white light-producing LED with correlated color temperature CCT of about 2700 K having a novel red phosphor having a total Sr content of 90% at the alkaline earth metal site (together with a green garnet phosphor) is distinctly higher here too, and an elevated color rendering index CRI is also achievable.

Red phosphors composed of the novel material system were subjected to a hydrolysis test in order to assess aging stability of the phosphor with respect to air humidity; see figure 8. Specifically, red phosphors composed of the $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ material system and one embodiment of the novel $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ phosphor of the invention were subjected to a hydrolysis test, in order to assess the aging stability of the phosphor with respect to air humidity. For this purpose, the phosphor powders were stored at 130°C and 100% rel. air humidity for 40 h. The relative absorption A of the phosphors in the blue spectral region (450 – 470 nm) as ordinate was measured both before and after this treatment. A measure of the stability of a phosphor to hydrolysis, i.e. the decomposition of the phosphor in the presence of water, is considered to be the decrease in the absorption capacity in the blue spectral region. With increasing Sr content, according to figure 8, for phosphors composed of the known $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ system, a significant increase in hydrolysis sensitivity is observed (solid rhombuses). Surprisingly, however, the novel $(\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu})$

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phosphor with $a = 0.8$ (corresponding in a formal sense to 90% Sr in a representation as $x = 0.9$) is more hydrolysis-stable (hollow rhombus) than a previously known $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ phosphor having an Sr content of 80% ($x = 0.8$).

In figure 9, the moisture stability of the novel phosphor is examined.

To improve the stability of the novel phosphor of the invention to hydrolysis, specimens of the phosphor were coated with an inert material (SiO_2) after the synthesis. Untreated and subsequently coated samples were subjected to a hydrolysis test, in order to assess the aging stability of the phosphor to air humidity. For this purpose, the phosphor powders were stored at 130°C and 100% rel. air humidity for 48-56 h. The quantum efficiency and the absorption of the phosphors in the blue spectral region (450 - 470 nm) were measured both before and after this treatment. A measure of the stability of a phosphor with respect to hydrolysis (decomposition of the phosphor in the presence of water) is considered to be the change in the relative conversion efficiency (calculated from the quantum efficiency and absorption in the spectral range of 450-470 nm) before and after the degradation test. The coating distinctly improves the stability.

Figures 10a and 10b show SEM images of various phosphors.

The figures show SEM images of the uncoated phosphor before and after the degradation process in different magnifications. What are shown are embodiments of the novel phosphor having the composition $\text{Sr}(\text{Sr}_{0.8}\text{Ca}_{0.2})\text{Si}_2\text{Al}_2\text{N}_6:1.2\%\text{Eu}$.

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Formation of cracks in the individual phosphor grains is apparent in the SEM images of the sample after the degradation test.

Figures 11a and 11b show SEM images of various phosphors.

The figures show SEM images of the coated phosphor before and after the degradation process in different magnifications.

What are shown are embodiments of the novel phosphor having the composition $\text{Sr}(\text{Sr}_{0.8}\text{Ca}_{0.2})\text{Si}_2\text{Al}_2\text{N}_6:1.2\%\text{Eu}$.

No formation of cracks in the phosphor grains is apparent in the SEM images of the sample after the degradation test.

Figure 12 shows thermal quenching characteristics of two red phosphors in comparison to one another. Both phosphors have a comparable emission color, with a dominant wavelength of approximately 600 nm. Surprisingly, the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ of the invention with $a = 0.8$ (solid squares), in spite of a higher Eu content (0.8%), has a smaller decrease in emission intensity I (ordinate) with rising temperature compared to the reference phosphor composed of the known $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ system; in that case, the Eu content is 0.6% (hollow rhombuses).

Figure 13 shows the relative external quantum efficiency QE for a previously known phosphor composed of the 1113-calsin system. The data are taken from EP 2 135 920.

What is reported therein about these phosphors composed of the $\text{CaAlSiN}_3:\text{Eu}$ system (referred to hereinafter as calsin) is that the conversion efficiency stagnates with rising activator content ($> 0.8\%$ Eu).

Similar behavior is also known for SCASN. The described phosphor composed of the $(\text{Sr},\text{Ca})\text{AlSiN}_3:\text{Eu}$ system with Sr

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content 80% shows strong dependence of the relative emission intensity of the luminescence signal on the activator content. This behavior is described, for example, in H. Watanabe et al., J. Electrochem. Soc., 2008, 155 (3), F31-F36. The behavior is shown in figure 14.

In contrast to pure calsin (cf. figure 13), the photoluminescence intensity actually collapses from a value of about 0.8% Eu or more for SCASN and only attains 60% of the maximum value.

Similar behavior is also described in US 8 274 215.

In the case of an Eu content of at least 1% ($x=0.01$), the luminescence signal decreases or virtually stagnates (Sr content: 80%). Figure 15 shows the figure derived therefrom (figure 15B therein). The inventors of US 8274 215 note in this regard that, with rising Eu content (up to the value of $x = 0.01$), the intensity of the photoluminescence increases, then it remains the same or decreases.

Figure 16 shows, in contrast, the relative emission intensity I as a function of the doping of Eu as activator, which replaces the alkaline earth metal content, for a novel phosphor. The Eu content is given here in %. Surprisingly, the novel phosphor (shown in the abscissa) shows behavior distinctly different from the prior art. With rising Eu content, the emission intensity I increases noticeably even in the case of an Eu content well above 1%, and in fact in an approximately linear manner. This property offers various technical advantages for the application. These include a relatively low phosphor requirement and the possibility of attaining color loci with relatively large x , understood as the first CIE component, and also high dominant wavelength λ_{dom} . With rising activator content Eu (shown as parameter y in %), the luminescence signal moves to greater

wavelengths in an approximately linear manner. This allows, for example, the color rendering index CRI of a warm white light-emitting LED to be increased; see also the other corresponding LED examples in the present application.

Figure 17 shows the influence of the degree of doping with Eu on the emission wavelength, shown as lambdadom (in nm). With rising activator content y for the novel phosphor, the luminescence signal moves to greater wavelengths in an approximately linear manner. This allows, for example, the color rendering index CRI of a warm white light-emitting LED to be increased; see also the other corresponding LED examples in the present application.

In order to determine the structure of the novel phosphor of the invention, crystals of the novel phosphor were chosen under a light microscope and prepared for a diffractometry study. The measurement was effected on a Bruker D8 Venture with rotating anode and CCD detector. A summary of the results (important goodness factors and the underlying refined parameters) can be found in the table in figure 21. The diffraction patterns collected were examined in great detail for their quenching conditions. A basic pattern discernible is a structure derived from AlN (wurtzite structure type) which can be described in the $P2_1$ space group.

The solution and refinement of the data set were effected with the JANA2006 software package (Petricek, V., Dusek, M. & Palatinus, L. (2006). Jana2006. The crystallographic computing system. Institute of Physics, Prague, Czech Republic.). The refinement proceeds very efficiently with the following restrictions: since Si and Al are indistinguishable by x-ray methods, all Si and Al positions were refined with the

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population of Si:Al = 1:1 as weighed out and a single thermal displacement factor for Si and Al.

In addition, all nitrogen atoms together and all alkaline earth metal atoms together were each described by one thermal displacement factor. All further parameters (e.g. the atomic position parameters) were freely refined.

Results of these single-crystal studies are discussed in detail hereinafter.

Figure 18 shows the crystallographic relationship between various light-emitting nitrides and AlN.

A whole series of known nitrides used as phosphors can be derived from AlN with wurtzite structure. Because of this fundamental structural relationship, the diffractograms (particularly x-ray powder diffractograms) of these compounds also often appear to be similar at first glance. However, differences are found in clearly pronounced details. The structures can differ significantly, as shown in figure 18 by way of example for the derivation of the unit cells for various compounds.

Figures 20a to 20c show, for the novel $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$ structure (the doping is unimportant for this fundamental consideration), why it clearly has to be described in the $P2_1$ space group and cannot be described in either of the two other space groups listed above.

According to figure 20, the single-crystal diffraction data are examined in reciprocal space for the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$. Figure 20a is a representation of the novel phosphor in reciprocal space viewed in the $[h01]$ direction.

The pseudohexagonal base structure is clearly apparent.

Figure 20b is a representation of the novel phosphor in reciprocal space viewed in the $[0kl]$ direction. The circled reflections are examples of reflections which cannot exist in

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the *Pna*21 space group. The occurrence thereof rules out description in this *Pna*21 space group because of the quenching conditions of various crystallographic space groups. The novel phosphor thus cannot have the same structure as, for example, MgSiN_2 , or MnSiN_2 .

Finally, figure 20c shows a representation of the novel phosphor in reciprocal space viewed in the $[h11]$ direction. The circled reflections are examples of reflections which cannot exist in the *Cmc*21 space group. The clearly apparent occurrence thereof rules out description in the *Cmc*21 space group. The novel phosphor thus cannot have the same structure as, for example, $(\text{Ca},\text{Sr})\text{AlSiN}_3$, LiSi_2N_3 , NaSi_2N_3 .

Figure 19 shows an overview of some structural data known from the literature for nitrides of similar composition (*Cmc*21 space group, NaSiO_3 structure type).

The single-crystal diffractometry solution of the novel structure $\text{Sr}(\text{Sr}_{\alpha}\text{Ca}_{1-\alpha})\text{Si}_2\text{Al}_2\text{N}_6$ is shown in detail in figure 21. This gives the lattice parameters, the unit cell, the radiation source used for the analysis, the reflections, the measured section of reciprocal space and further data.

Figure 22 gives the interatomic distances in the novel structure $\text{Sr}(\text{Sr}_{\alpha}\text{Ca}_{1-\alpha})\text{Si}_2\text{Al}_2\text{N}_6$ in detail. In a direct comparison with nitrides of similar composition, for example SrAlSiN_3 (ICSD 419410), CaAlSiN_3 (ICSD 161796) or $(\text{Sr},\text{Ca})\text{AlSiN}_3$ (ICSD 163203) (cf. table in figure 19 for further examples), it should be noted that there is a somewhat larger and a somewhat smaller environment around the alkaline earth metal atoms Sr and Ca. In SrAlSiN_3 , CaAlSiN_3 and $(\text{Sr},\text{Ca})\text{AlSiN}_3$, for the alkaline earth metal atoms, there

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is only a pentacoordinated position with a mean Sr-N distance of 267 pm. In the novel structure $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$ of the invention, Sr1 forms a hexacoordinated environment with a mean Sr1-N distance of 272 pm; Sr2/Ca2 forms a pentacoordination with a mean Sr2/Ca2-N distance of 264 pm.

Figure 23 compares the crystallographic data and positional parameters (according to single crystallography) for a novel compound $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$ with the corresponding data for the previously known $\text{SrAlSiN}_3:\text{Eu}$ (on the right). There is a distinct difference between the crystal systems and space groups.

Figure 24 illustrates the structure of the novel phosphor type $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$. Figure 24a shows a view of the layers of $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$. The layers are derived from AlN. Compared to AlN, individual tetrahedra are absent and are replaced by an alkaline earth metal ion. The tetrahedra are distinctly distorted compared to AlN. However, all bond lengths and angles are similar to other nitridosilicates. Figure 24b shows the phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$ of the invention from the [010] direction. The 3D network of the $(\text{Si}/\text{Al})\text{N}_4$ tetrahedra is clearly apparent. Running in the $a-c$ plane are layers that are linked in the b direction (not shown) to form a network. Intercalated between these, layer by layer in each case, are the pure Sr position (shown as white circles) and the position having a mixed Sr/Ca population (shown as black circles).

Figure 24c shows, for comparison, the structure, known from the literature, of $(\text{Sr}_{0.846}\text{Ca}_{0.211})\text{AlSiN}_3$ (ICSD 163203) from the [010] direction. Here, all the Sr/Ca positions (black) have mixed populations. There are no pure Sr positions.

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This ordering into a position having a mixed Sr/Ca population and a position fully populated by Sr alone in the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ of the invention is advantageous, for example over the structure of SCASN (cf. figure 24c), where only a position having a mixed population is available for the activator atoms (doping), which leads to broadening of the emission, this of course being based on the interaction between the activator and the surrounding host lattice, and to stronger quenching properties. The $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ structure of the invention, by contrast, offers the activator, preferably Eu here, an ordered Sr position, without disorder and without the associated disadvantages. The improved properties of luminescence can be explained plausibly by this structure. The Eu, according to this model concept, populates predominantly the pure Sr plane only, and the mixed plane to a lesser degree.

Proceeding from figure 24b, it is also possible to represent a phosphor having the lower symmetry according to space groups 1 to 3 of the International Tables Cryst. A, i.e. space groups $\text{P}1$, $\text{P}\bar{1}$, $\text{P}2$, in that, for example, the mixed layer has been split up into planes having different population (in part) by means of pure Sr alongside a mixed population.

In figure 25 and figure 26 there is a crystallographic evaluation. Figure 25 shows a Rietveld refinement of the x-ray powder diffractogram of specimen TF162bG/12, an embodiment of the novel phosphor of the invention having the weighed-out composition $\text{Sr}(\text{Sr}_{0.8}\text{Ca}_{0.2})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$. The diffractogram can be very well described with the structure

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model found by the single-crystal x-ray structure analysis for $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{N}_2\text{N}_6$ (R_{profil} 7%, R_{Bragg} 6%).

Figure 26 shows an enlarged section of the Rietveld refinement of the x-ray powder diffractogram of TF162bG/12. The reflection marked with an arrow is an example of a reflection of $\text{Sr}(\text{Sr}_{0.8}\text{Ca}_{0.2})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ which can occur in the $P2_1$ space group and other lower-symmetry space groups such as the $P1$ space group. The occurrence of this reflection definitely rules out description of the novel phosphor in the higher-symmetry $Cmc21$ and $Pna21$ space groups of the other AlN-related nitridosilicatic phosphors.

Figure 27/28 and figure 29/30 and 31/32 and figure 33/34 show, by way of example, further Rietveld refinements of the x-ray powder diffractogram of other embodiments, as shown in each diagram, each with a characteristic section.

Figures 35a and b show the absorption and emission characteristics of a phosphor of the invention compared to known phosphors.

Figure 35a shows emission spectra and figure 35b reflectance spectra of an embodiment of the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a = 0.8$ and 0.8% Eu compared to known phosphors composed of the $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ system with comparable dominant wavelength λ_{dom} ($\lambda_{\text{dom}} \approx 600$ nm). The wavelength λ is plotted against the intensity I and the reflectance R . The emission spectra show an unexpectedly narrow spectral emission of the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a = 0.8$. At the same time, the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a = 0.8$ features a

strong absorption; see figure 35b. The absorption is found to be approximately 1-R.

Figures 36a and b show the absorption and emission characteristics of a further phosphor of the invention compared to known phosphors.

Figure 36a shows emission spectra and figure 36b reflectance spectra of an embodiment of the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a = 0.8$ and 1.2% Eu compared to known phosphors composed of the $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ system with comparable dominant wavelength I_{dom} ($\text{I}_{\text{dom}} \approx 602-603 \text{ nm}$). The wavelength λ is plotted against the intensity I and the reflectance R . The emission spectra show an unexpectedly narrow spectral emission of the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a = 0.8$. At the same time, the novel phosphor $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a = 0.8$ features a strong absorption; see figure 36b. The absorption is found to be approximately 1-R.

The novel phosphor described here offers the following advantages in particular:

- lower half-height width of emission, associated with higher luminous efficiency at the same dominant wavelength,
- the possibility of achieving higher activator concentrations of Eu at > 0.8% with simultaneously high quantum efficiency and conversion efficiency, associated with a smaller phosphor demand in LED applications and simplified processibility,
- improved aging stability with respect to moisture compared to conventional $(\text{Sr},\text{Ca})\text{AlSiN}_3:\text{Eu}$ having low Sr content, and
- improvement of thermal stability.

Figure 37 shows a semiconductor component which serves as a light source (LED) for white light.

For use in a white LED together with a GaInN chip, for example, a construction similar to that described in US 5 998 925 is used. The structure of such a light source for white light is shown explicitly in figure 37. The light source, which is the radiation source for the primary radiation, is a semiconductor component which can emit primary radiation in the UV and/or blue wavelength range. For example, the radiation source used may be a semiconductor component (chip 1) of the InGaN type with a peak emission wavelength of 460 nm having a first and second electrical connection 2, 3, embedded into an opaque base housing 8 in the region of a recess 9. One of the connections 3 is connected to the chip 1 via a bonding wire 14. The recess has a wall 17 which serves as reflector for the blue primary radiation from the chip 1. The recess 9 is filled with a potting compound 5 comprising, as main constituents, a potting compound and phosphor pigments 6 (less than 50% by weight). Further small proportions are accounted for, for example, by methyl ether and Aerosil inter alia. The phosphor pigments are a mixture of various phosphor mixtures described here, which also contain phosphors of the invention, for example LuAG:Ce pigments and pigments of the novel phosphor.

Generally, in the case of an LED chip that emits UV radiation as radiation source, it is possible to use a phosphor mixture composed of at least three different phosphors (blue-emitting phosphor, for example BaMgAl₁₀O₁₇:Eu²⁺ or (Ba,Sr,Ca)₅(PO₄)₃Cl:Eu²⁺, together with a green/yellow-emitting phosphor, for example one of the garnet phosphors

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described here, and an orange/red-emitting phosphor, for example one of the invention), and, in the case of a blue light-emitting LED chip as radiation source, to use a phosphor mixture composed of at least two different phosphors (a green/yellow phosphor, for example one of the garnet phosphors described here, and an orange/red-emitting phosphor, for example one of the invention).

Figure 38 shows a section from an area light 20 as lighting unit. It consists of a common carrier 21, onto which has been bonded a cuboidal outer housing 22. The upper side thereof has been provided with a common cover 23. The cuboidal housing has recesses in which individual semiconductor components 24 are accommodated. They are UV light-emitting diodes and/or light-emitting diodes that emit blue light as primary radiation with a peak emission of 380 nm. The conversion to white light is effected by means of conversion layers that are directly within the casting resin of the individual LEDs, similarly to the manner described in figure 37, or layers 25 applied to all the surfaces accessible to the UV radiation (these are especially ceramic surfaces or plates). These include the inner surfaces of the side walls of the housing, the cover and the base section. The conversion layers 25 comprise, in the case of UV light-emitting LEDs, three phosphors which emit in the red-orange, yellow-green and blue spectral regions with utilization of phosphors of the invention. When the LEDs as radiation sources emit blue radiation as primary radiation, as already described above, it is also possible for only two different phosphors that emit in the green-yellow or orange-red to be present in the phosphor mixtures.

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In a number of further embodiments of the present invention, conventional blue light-emitting InGaN LEDs were provided with a conventional silicone potting material with different amounts of various phosphors of the invention or conventional phosphors as comparative examples embedded therein. In this case, phosphors of the invention in particular are to be used partly together with other phosphors for color conversion of blue primary light to the red or yellow or yellow-orange wavelength range. These phosphors may also be used, inter alia, for full conversion of the primary light from the primary radiation source to the red or yellow or yellow-orange wavelength range.

Alternatively, the primary radiation source used, rather than an LED such as an InGaN LED, may also be an organic light-emitting device (OLED) comprising a layer stack of organic semiconductive layers disposed between an anode and a cathode. In this case, at least one of the electrodes has to be transparent to the radiation produced by the OLED, in which case the phosphors of the invention can be disposed in the beam path of the radiation above the transparent electrode.

Embodiments for full conversion of red:

The table in figure 40a shows the dominant wavelength of the blue-emitting LED ($\lambda_{\text{dom(blue LED)}}$), the chemical formulae of the inventive and conventional phosphors used, and the concentrations thereof in the potting material (percent by weight based on the overall potting material), the x and y color coordinates of the secondary radiation converted in the CIE color space and the resulting luminous flux $\Phi_{v(\text{potting})}$ and

the radiant power $\Phi_{e(potting)}$ of the potted LEDs, in each case relative to the $\Phi_{e(no\ potting)}$ value for the unpotted LEDs (figures relative to the comparative example in each case). All further tables likewise include this parameter and in some cases further parameters, for example mixing ratios in the case of mixing of two different phosphors in the potting material of an LED.

It can be inferred unambiguously from this table in figure 40a that, with a very similar color locus in the red-orange region (x and y CIE color coordinates), the phosphor of the invention according to inventive example 1 compared to the conventional phosphor according to comparative example 1 has a higher luminous flux and also a higher radiant power. This can also be inferred from the corresponding emission spectrum of figure 40b, in which it is clearly apparent that inventive example 1 has a higher light intensity compared to the LED of comparative example 1.

In the table in figure 41a, again, the luminous flux and the radiant power of a blue-emitting LED having a conventional CaAlSiN_3 phosphor (comparative example 2, phosphor effected according to EP patent application EP 1 696 016 A1) are compared with an LED with one embodiment of a phosphor of the invention incorporated into the silicone potting material thereof (inventive example 2). Because of the good stability of the phosphors of the invention, they can also be used with a higher proportion of activator dopants (in the present case 5%) compared to CaAlSiN_3 phosphors, where a corresponding color locus can be achieved only with very low Eu dopings (in the present case 0.4%). It can again be inferred from the table in figure 41a that, with comparable color loci of the two LEDs in the red color space, the LED having the phosphor

of the invention has a higher luminous flux and a higher radiant power. This can likewise be inferred from the emission spectrum of figure 41b, in which the emission spectra of comparative example 2 and of inventive example 2 are compared.

The second phosphors described in EP patent application EP 1 696 016 A1 are especially phosphors including the elements M, A, D, E, and X, where M is one or more elements selected from the group consisting of: Mn, Ce, Pr, Nd, Sm, Eu, Tb, Dy, Ho, Er, Tm, and Yb, A is one or more elements selected from the group consisting of divalent metal elements other than M, D is one or more elements selected from the group consisting of tetravalent metal elements, E is one or more elements selected from the group consisting of trivalent metal elements, X is one or more elements selected from the group consisting of: O, N, and F, and where the phosphors have the same crystal structure as CaAlSiN_3 .

In addition, second phosphors used may also be phosphors of the general formula $(\text{M}_{1-x}\text{Eu})_x(\text{Al},\text{Q})(\text{Si},\text{Y})\text{N}_3$ where M = Ca, Sr, alone or in combination or in combination with other divalent and/or monovalent elements, for example Li, Q = trivalent cation other than Al^{3+} , Y = other tetravalent cations other than Si^{4+} , N^{3-} may be partly replaced by O^{2-} , F^- , Cl^- , Br^- , C^{4-} .

In the table in figure 42a, comparative example 3 and inventive example 3 are compared. In inventive example 3 again, a conventional CaAlSiN_3 phosphor $\text{CaAlSi}(\text{N},\text{O})_3:\text{Eu}$ (0.4%) is used as the first phosphor, with additional incorporation of a further embodiment of a phosphor of the invention in the potting material as a second phosphor. In the present case, the CaAlSiN_3 phosphor is capable of

absorbing the short-wave components of the secondary radiation converted by the phosphor of the invention and converting it to red light with longer wavelengths compared to the starting light. This approach has the advantage that, in contrast to wavelength-specific filters, the radiation is not just absorbed but also emitted again because of the conversion in the phosphor, which leads to an increased radiant power of the LED. It can again be inferred from the table that, with a similar color locus in the CIE color space, both the luminous flux and the radiant power of inventive example 3 have greatly increased compared to comparative example 3. This can likewise be inferred from the emission spectrum of figure 42b.

Analogously to inventive example 3 in inventive example 4 as well, which is compared with a comparative example 4 in the table in figure 43a, a conventional CaAlSiN_3 phosphor is used, which absorbs the short-wave components of the light converted by the phosphor of the invention and emits them again as red light of a higher wavelength. This table too, and the corresponding emission spectrum in figure 43b, show that, given similar color loci, the LED comprising the phosphor of the invention has a higher luminous flux and a higher radiant power compared to the LED comprising the conventional phosphors.

Two inventive examples 5 and 6 are compared in the table in figure 44a to a comparative example 5. In all LEDs, the blue primary radiation is converted to an orange secondary radiation, with exclusive use of conventional phosphors in comparative example 5 and of different embodiments of phosphors of the invention with a cerium-activated yttrium aluminum garnet phosphor in each of inventive examples 5 and

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6. It can likewise be inferred from this table, just as from the corresponding emission spectrum of figure 44b, that the LEDs having the combination of phosphors including phosphors of the invention have a higher luminous flux and a higher radiant power than LEDs comprising conventional phosphors.

Further lighting devices of the invention are to be described hereinafter, where different amounts of phosphors of the invention having high europium concentrations are present as dopant in the potting material of the radiation source that emits the primary radiation, for example a blue LED. Lighting devices of this kind can also be used, for example, for full conversion of the primary radiation from the radiation source to a secondary radiation, for example in a red or yellow or yellow-orange wavelength range. More particularly, it is possible for lighting devices of the invention, in particular embodiments of the present invention, to have, as a radiation source that emits primary light, a blue LED having a dominant wavelength of 300-500 nm, preferably 400-500 nm, further preferably 420-470 nm, and a phosphor of the invention having the general formula $Sr(Sr_aM_{1-a})Si_2Al_2N_6:D$ where M is selected from the group of Ca, Ba, Zn, Mg, preferably Ca, and D is preferably Eu, and where the europium concentrations may be ≥ 6 mol%, further preferably ≥ 8 mol%. The radiation emitted by these radiation sources may have a half-height width FWHM of ≤ 90 nm, preferably ≤ 85 nm, and a dominant wavelength of ≥ 607 nm, preferably ≥ 609 nm.

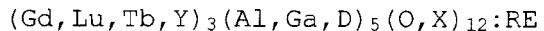
The table in figure 45a shows various embodiments of lighting devices of the invention in which different concentrations of phosphors of the invention have been incorporated in the potting material of a conventional blue InGaN LED with a standard silicone potting material. In spite of the high

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europtium concentrations, figure 45b shows that the half-height widths of the radiation emitted increase to a lesser degree with increasing europtium concentrations in the case of phosphors of the invention having the general formula $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ than in the case of conventional phosphors of the formulae $\text{CaAlSiN}_3:\text{Eu}$ and $\text{Sr}_2\text{Si}_5\text{N}_8:\text{Eu}$. Figure 45c likewise shows that, in the case of phosphors of the invention having an increasing europtium content, surprisingly, the quantum efficiency decreases less significantly compared to conventional phosphors. Figure 45d shows the emission spectra of the three inventive examples 1 to 3, it being clearly apparent that, apart from a very small proportion of the primary radiation, the entire radiation emitted by the lighting device is attributable to converted secondary radiation having a low half-height width FWHM. These properties of phosphors of the invention allow the provision of lighting devices which emit deep red light for the purposes of full conversion of the primary radiation.

The phosphors of the invention may be used with a multitude of different garnet phosphors as second phosphors. These may especially have the general structural formula:

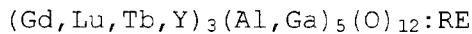


with X = halide, N or divalent element, D = tri- or tetravalent element and RE = rare earth metals as activator, especially cerium with an optional co-dopant, for example lanthanoids, e.g. Pr, Sm, Nd.

The garnets may additionally also have the following general formula:

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with RE = rare earth metals as activator, especially cerium with an optional co-dopant, for example lanthanoids, e.g. Pr, Sm, Nd.

Embodiments of streetlighting applications:

In a further embodiment of the present invention, it is possible to provide lighting devices of the invention which can especially also be used for general lighting applications, for example streetlighting, with a CRI ≥ 70 and high color temperatures (~ 5000 K).

More particularly, these lighting devices may have, as radiation source, a blue LED having a dominant wavelength of 300-500 nm, preferably 400-500 nm, further preferably 420-470 nm, with at least one phosphor of the invention having the general formula $Sr(Sr_aM_{1-a})Si_2Al_2N_6:D$ where M is selected from the group of Ca, Ba, Zn, Mg, with D = Eu, present in the beam path of the radiation source as first phosphor and at least one yellow/green-emitting garnet phosphor of the general formula $(Y, Lu, Gd)_3(Al, Ga)_5O_{12}:Ce$ present as second phosphor.

The first phosphor used here may be a phosphor of the invention having the general formula $Sr(Sr_aM_{1-a})Si_2Al_2N_6:D$ where M is selected from the group of Ca, Ba, Zn, Mg, preferably Ca, and D is preferably Eu, with $a \geq 0.8$, preferably $a \geq 0.82$, further preferably $a \geq 0.85$, and a europium content of 0.1-5 mol%, preferably 0.1-3 mol%, further preferably 0.1 to 2 mol%. The phosphor of the invention has a peak emission wavelength of about 600-640 nm,

preferably 605-625 nm, and a half-height width FWHM of < 85 nm, preferably < 80 nm, further preferably < 78 nm. Such a phosphor of the invention together with the garnet phosphor as second phosphor can give a lighting device in which a wide range of correlated color temperature CCT within the range of at least 6500-4000 K, preferably 6500-3000 K, is possible, where the CRI is at least 70.

In this case, the garnet phosphor as second phosphor may especially have the general formula $\text{Lu}_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ or $(\text{Y},\text{Lu})_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$, in which case the maximum excitability is preferably between 440-455 nm, further preferably between 454-450 nm. The yellow/green-emitting garnet phosphor is selected such that it has a high conversion efficiency and high thermal stability. A preferred yellow/green-emitting phosphor is a $(\text{Y},\text{Lu})_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ having a cerium content of 0.5-5 mol%, preferably 1-3 mol%, and a Y content of 0-50 mol%, preferably 0-30 mol%, so as to result in a phosphor of the general formula $(\text{Lu}_{1-x}\text{Y}_x)_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ with $x = 0$ to 0.5, preferably $x = 0$ to 0.3. Other variants of the garnet phosphor are also possible, having similar spectral characteristics, especially variants with $(\text{Y},\text{Lu})_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ where at least some of the Al has been exchanged for Ga.

In a further embodiment, a garnet phosphor of the following general formula $\text{Y}_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ is used, having a maximum excitability in the range of 440-455 nm, preferably 445-450 nm. The preferred yellow/green-emitting phosphor is selected such that it has a high conversion efficiency and high thermal stability. The preferred yellow/green-emitting phosphor is a phosphor of the general formula $\text{Y}_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ having a cerium content of 1.5-5 mol%, preferably 2-5 mol%,

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and a gallium content of 0-50 mol%, preferably 0-30 mol%. Within this system, however, other element combinations are also possible, which result in similar spectral properties.

Hereinafter the optical properties of two comparative examples 1 and 2 where a garnet phosphor has been combined with a conventional 2-5-8 phosphor in the standard silicone potting material of a conventional InGaN LED having the dominant wavelength of 444.5 or 444.6 nm at 350 mA are to be compared with three embodiments of the present invention. The area of each of the LED chips is 1 mm².

The compositions of the various phosphors and the respective concentrations of the phosphors in the standard potting material and the relative proportions of the two phosphors are given in figure 46a.

It can be inferred from the table in figure 46b that, given similar color coordinates, the LEDs of inventive examples 1-4 comprising the phosphors of the invention, at a driver current of 350 mA, have a higher or similar conversion efficiency as the ratio of the luminous flux Φ_v of an LED having potting material filled with the phosphor mixtures and the radiant power Φ_e of an LED having a clear potting material without phosphors $\Phi_v(\text{filled potting})/\Phi_e(\text{clear potting})$ compared to the LEDs comprising the conventional phosphor mixture, but the CRI is higher. The color point correction was conducted by taking a theoretical model of comparative example 2 and extrapolating the LED efficiency of the comparative example for the color points of the other inventive examples. The conversion efficiencies were each stated as the relative conversion efficiencies in relation to inventive example 1.

Figure 46c shows the same measurement data as figure 46b, but at a driver current of 1000 mA. In that figure too, the inventive examples again have a higher CRI.

In summary, it can be stated that neither of comparative examples 1 and 2 attains the necessary CRI of 70 at 3000 K for streetlighting applications. For this reason, comparative examples 1 and 2 can be used either only in lighting devices having a correlated color temperature CCT between 6500-4000 K and not within a range of 6500-3000 K, or an additional, third phosphor has to be used to improve the color temperature, but this complicates the manufacturing process for the lighting devices. Thus, lighting devices of the invention, because of the phosphor of the invention, can have a sufficient CRI > 70 within a broad color temperature range of 6500-3000 K for streetlighting applications. In contrast to conventional phosphor mixtures, no additional, third phosphor is needed for the achievement of a broad color temperature range.

Figures 46d and 46e show the normalized LED spectra of inventive examples 1 and 2 and of the corresponding comparative examples 1 and 2 at a driver current of 350 mA (figure 46d) and the normalized LED spectra of inventive examples 3 and 4 and of comparative examples 1 and 2 at a driver current of 350 mA (figure 46e).

In further embodiments of the present invention, the efficiency of the lighting devices, for example for streetlighting, at a particular color temperature CCT and a particular CRI, can be improved by adding a further, third phosphor to the phosphor mixtures. More particularly, a

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lighting device in this embodiment of the present invention may have a radiation source having a blue light-emitting LED having a dominant wavelength between 440-455 nm, and also a red-emitting phosphor of the invention having a peak wavelength between 605-620 nm, preferably 605-616 nm, and a half-height width FWHM \leq 80 nm, preferably \leq 78 nm, as the first phosphor, a green/yellow-emitting phosphor having a peak wavelength between 540-565 nm, preferably 546-556 nm, and a half-height width FWHM \geq 100 nm, preferably \geq 105 nm, as the second phosphor, and a yellow/orange-emitting phosphor having a peak wavelength of 580-590 nm, preferably 582-588 nm, and a half-height width FWHM \leq 80 nm, preferably \leq 78 nm, as the third phosphor.

The first phosphor used may especially be a red-emitting phosphor of the general formula $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{D}$ where M is selected from the group of Ca, Ba, Zn, Mg, preferably Ca, having a peak emission between 605-620 nm, preferably 605-616 nm, and a half-height width FWHM below or equal to 80 nm, preferably below or equal to 78 nm. It may especially be the case here that the value $a \geq 0.8$, preferably $a \geq 0.84$, and the europium content is between 0.1-5 mol%, preferably 0.1-3 mol%, further preferably between 0.1-2 mol%. These red-emitting phosphors feature a high thermal stability and a high conversion efficiency under operating conditions typical for streetlighting.

More particularly, the green/yellow-emitting phosphor as the second phosphor may be a garnet phosphor of the general formula $(\text{Y}, \text{Lu}, \text{Gd}, \text{Tb})_3(\text{Al}, \text{Ga})_5\text{O}_{12}:\text{Ce}$ which is matched to the emission wavelengths of the blue LED and the two other phosphors. More particularly, the phosphor may be a green-emitting garnet phosphor of the general formula

$(Y, Lu)_3(Al, Ga)_5O_{12}:Ce$, especially a yellow/green-emitting phosphor of the general formula $(Y, Lu)_3(Al, Ga)_5O_{12}:Ce$ having a cerium content of 1-5 mol%, preferably 2-4 mol%, and an yttrium content of 0-50 mol%, preferably 0-30 mol%, further preferably 0-20 mol%, and a gallium content of 0-50 mol%, preferably 0-30 mol%, further preferably 0-15 mol%, and so the result is a garnet phosphor of the following general formula: $(Lu_{1-x} Y_x)_3(Al_{1-y}Ga_y)_5O_{12}:Ce$ with $x = 0$ to 0.5, preferably $x = 0$ to 0.3, further preferably $x = 0$ to 0.15 and $y = 0$ to 0.5, preferably $y = 0$ to 0.2, further preferably $y = 0$ to 0.15. Other element combinations within the general formula $(Y, Lu, Gd, Tb)_3(Al, Ga)_5O_{12}:RE$ with RE = rare earth metal, e.g. Ce, are likewise possible.

The third phosphor used may especially be a phosphor of the general formula $(Ca, Sr, Ba)_2(Si, Al)_5(N, O)_8:Eu$ which, together with the primary light source, for example a blue LED, and the green/yellow-emitting phosphor and the red-emitting phosphor of the invention, achieves a CRI ≥ 65 , preferably ≥ 70 , over a wide range of correlated color temperature CCT (at least of 6500-4000 K, further preferably of 6500-3000 K). In order to achieve the above-described spectral properties for the yellow/red- or yellow/orange-emitting phosphor as the third phosphor, this has a europium content of 0.1-5 mol%, preferably of 0.1-3 mol%, further preferably of 0.1-2 mol%, and a barium content of 50-100 mol%, preferably 70-100 mol%, further preferably 80-100 mol%, and a calcium content of 0-20 mol%, preferably 0-10 mol%, where the proportion of strontium is chosen such that the alkaline earth metals barium, strontium and calcium together with the europium dopant add up to 100%.

There follows a description of one embodiment of a lighting device of the invention having three different first to third phosphors in relation to a comparative example. Comparative example 1 is a lighting device having a blue LED chip (dominant wavelength 445 nm) having a chip area of 1 mm², in which 14% by weight of a phosphor mixture composed of two different phosphors is present in the standard silicone potting material of the LED, where the ratio of the green to the red phosphor is 4.7:1.

The green-emitting phosphor here is a phosphor of the formula $(Lu_{0.85}Y_{0.15})_3Al_5O_{12}:Ce$ (3 mol%), and the red-emitting phosphor a phosphor of the formula $(Sr_{0.5}Ba_{0.5})_2Si_5N_8:Eu$ (1 mol%), the LED being operated at a driver current of 350 mA. In inventive example 1, present in the standard silicone potting material of a blue LED having a dominant wavelength of 445 nm at a driver current of 350 mA is 14% by weight of a phosphor mixture comprising the following three first to third phosphors: $Sr(Sr_{0.86}Ca_{0.14})Si_2Al_2N_6:Eu$ (0.8 mol%) as the first phosphor, $(Lu_{0.85}Y_{0.15})_3Al_5O_{12}:Ce$ (3 mol%) as the second phosphor and $(Sr_{0.1}Ba_{0.9})_2Si_5N_8:Eu$ (1 mol%) as the third phosphor, where the ratio of first phosphor:second phosphor:third phosphor is 0.67:5.3:0.33. The area of the LED chip is again 1 mm².

Figure 47a shows, in tabular form, a list of the CIE color coordinates and of the CRI and the conversion efficiency as the ratio of the luminous flux Φ_v of an LED having a potting material filled with the phosphor mixtures and the radiant power Φ_e of an LED having a clear potting material without phosphors $\Phi_{(v(filled\ potting))}/\Phi_{e(clear\ potting)}$ and the luminous efficiency for comparative example 1 and inventive example 1 of the present invention at a driver current of 350 mA and a dominant wavelength of 444.6 nm. It is apparent that both the

luminous efficiency and the conversion efficiency in the inventive example are higher than in the comparative example.

Figure 47b shows further results for inventive example 1 and for comparative example 1 at a driver current of 350 mA at different temperatures of 25°C and 85°C. Here too, it is apparent that the luminous efficiency is higher in inventive example 1 than in comparative example 1.

Figure 47c shows the emission spectra of comparative example 1 and of inventive example 1 at 25°C at a color temperature of 4000 K and a driver current of 350 mA. The two examples have roughly comparable emission spectra.

Figure 47d shows the spectral efficiency (LER, lm/W_{opt}) of inventive example 1 compared to comparative example 1 as a function of the dominant wavelength of the LED chip at a color temperature of 4000 K based on ray tracing simulations. It is clearly apparent that, in the case of the phosphor mixture of the invention, composed of three phosphors, the spectral efficiency is greater than in the comparative example. The data shown in the figures 74e to 47i that follow are also based on ray tracing simulations, with selection of color loci on the Planckian locus with the CCT specified for the simulations.

Figure 47e shows that the range for the CRI with the phosphor mixture of the invention, comprising three phosphors, can be adjusted over a very wide range between 53 and 76 over a range of the correlated color temperature CCT of 3000-6500 K (see areas shaded gray). The dominant wavelength of the LED chip is 448 nm and the dashed and dotted lines in this figure show the CRI for two comparative examples. Inventive example

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1 comprises a mixture of three different phosphors of the following composition: $(Lu_{0.85}Y_{0.15})_3Al_5O_{12}:Ce$ (3 mol%), $(Sr_{0.1}Ba_{0.9})_2Si_5N_8:Eu$ (1 mol%) and $Sr(Sr_{0.86}Ca_{0.14})Si_2Al_2N_6:Eu$ (0.8 mol%), whereas comparative example 1 contains the two following phosphors: $(Lu_{0.8}Y_{0.2})_3Al_5O_{12}:Ce$ (4.5 mol%) and $(Ca_{0.025}Sr_{0.475}Ba_{0.5})_2Si_5N_8:Eu$ (2.5 mol%), and comparative example 2 comprises the following phosphors: $(Lu_{0.85}Y_{0.15})_3Al_5O_{12}:Ce$ (3 mol%) and $(Sr_{0.5}Ba_{0.5})_2Si_5N_8:Eu$ (1 mol%).

Figures 47f to 47i show the luminous efficiency LER for this inventive example 1 and comparative examples 1 and 2 as a function of the proportion of the red phosphor of the invention for inventive example 1 at various CCTs (figure 47f: 3000 K on Planck; figure 47g: 4000 K on Planck; figure 47h: 5000 K on Planck and figure 47i: 6500 K on Planck). For comparative examples 1 and 2 which contain only two phosphors, the luminous efficiency LER and the proportion of the red phosphor are already determined by the desired color point/CCT. Therefore, the values on the x axis relate only to inventive example 1, and the values shown there give the proportion of the red phosphor in relation to the orange phosphor in the phosphor mixture of the invention containing the three phosphors. The mixture of the red and green phosphors is then mixed with the green-yellow phosphor, in order to achieve the desired color locus (LER at CRI 70 marked with an arrow for embodiments). It can additionally be inferred from these figures that, at a CRI of 70, the phosphor mixture of the invention has an LER higher than in the case of conventional phosphor mixtures. Particularly in the case of high correlated color temperatures CCT, the proportion of the red phosphor in the mixture can be gradually reduced and, correspondingly, the proportion of the orange-emitting phosphor increased, since a high proportion

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of red-emitting phosphor is required particularly at low CCT. Zero or only small proportions of deep red emission in the LED spectrum generally have a positive effect on the LED efficiency.

Embodiments of backlighting applications:

A further embodiment of the present invention is directed to phosphor mixtures for backlighting applications. For backlighting applications, a broad color space has to be achieved with narrow-band red- and green-emitting phosphors, the phosphor mixture determining the optical properties of the LED, such as brightness, efficiency and robustness.

For the sRGB/Rec709 color space, lighting devices which can be used as backlighting devices for LCDs, for example, are especially those having, as primary radiation source, a blue LED having a dominant emission wavelength of 430-470 nm, preferably 440-460 nm, further preferably 445-455 nm, and also containing a green/yellow-emitting garnet phosphor of the general formula $(\text{Lu},\text{Y},\text{Gd},\text{Tb})_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ and also a red-emitting phosphor, especially an inventive phosphor of the general formula $\text{Sr}(\text{Sr}_{a}\text{M}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{D}$ where M is selected from the group of Ca, Ba, Zn, Mg, preferably Ca, and D is preferably Eu. It is optionally also possible for further converters or phosphors or non-converting materials such as diffusers to be present in the phosphor mixture.

The garnet phosphor may especially have the general composition $(\text{Lu},\text{Y})_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ and may also take the form of yttrium aluminum gallium garnet having a gallium content of $20 \text{ mol\%} \leq x \leq 60 \text{ mol\%}$, further preferably $30 \text{ mol\%} \leq x \leq 50 \text{ mol\%}$, further preferably $30 \text{ mol\%} \leq x \leq 45 \text{ mol\%}$, so as to

result in the general formula: $Y_3(Al_{1-x}Ga_x)_5O_{12}:Ce$ with $0.2 \leq x \leq 0.6$, further preferably $0.3 \leq x \leq 0.5$, further preferably $0.3 \leq x \leq 0.45$. In addition, the garnet phosphor may also take the form of lutetium aluminum gallium garnet having the following general formula: $Lu_3(Al_{1-x}Ga_x)_5O_{12}:Ce$ with $0 \leq x \leq 0.6$, further preferably $0 \leq x \leq 0.4$, further preferably $0 \leq x \leq 0.25$, with a cerium content of 0.5-5 mol%, preferably 0.5-3 mol%, further preferably 0.5-2.5 mol%, based in each case on the rare earth metals.

The red-emitting phosphor of the invention may have an activator content of ≥ 2 mol%, preferably an activator content of ≥ 3 mol%, further preferably a content of ≥ 4 mol%, where the divalent metals, which are preferably strontium and calcium, have a calcium content of ≤ 15 mol%, preferably ≤ 10 mol%, further preferably ≤ 8 mol%, so as to result in the following general formula:

$Sr(Sr_aCa_{1-a})Si_2Al_2N_6:Eu$ with $a \geq 0.7$, preferably $a \geq 0.8$, further preferably $a \geq 0.84$.

Both the abovementioned garnet phosphors may be used, the use of the yttrium aluminum gallium garnet rather than the lutetium aluminum gallium garnet bringing the advantage that the yttrium garnet has a lower specific density, such that less phosphor is required for the phosphor mixtures and, at the same time, a lower percentage of rare earth metals has to be used for the production of the phosphors, and so they can be produced less expensively.

For backlighting applications having elevated demands on the color space (for example Adobe RGB, NTSC or DCI-P3), phosphor mixtures having very narrow-band-emitting green-yellow phosphors are required. Preferably, lighting devices of this

kind have, as radiation source, a blue LED having a dominant wavelength between 430-470 nm, preferably 440-460 nm, further preferably 445-455 nm.

The green/yellow-emitting phosphors used may be nitridoorthosilicates which have the general composition $AE_{2-x}RE_xSiO_{4-x}N_x:Eu$ with AE = Sr, Ca, Ba, Mg, RE = rare earth metals and/or $AE_{2-x}RE_xSi_{1-y}O_{4-x-2y}N_x:Eu$, and AE and RE as defined in the previous example, and which are more particularly described in the patent application WO 2011/160944. It is likewise possible to use orthosilicates of the general formula $AE_2SiO_4:Eu$ with AE = Ca, Ba, Mg, Sr. Both the nitridoorthosilicates and orthosilicates having the abovementioned empirical formulae preferably include a combination of at least two alkaline earth metals, further preferably a combination of strontium and barium having the ratio of $0.5 \leq Ba:Sr \leq 2$, further preferably $0.75 \leq Ba:Sr \leq 1.25$. The nitridoorthosilicates may also be described by the general formula $AE_{2-x}L_xSiO_{4-x}N_x:RE$ where AE contains one or more elements selected from Mg, Ca, Sr, Ba and RE contains one or more elements selected from rare earth metals, preferably at least Eu, and L contains one or more elements selected from rare earth metals other than RE, with $0 < x \leq 0.1$, preferably $0.003 \leq x \leq 0.02$. A further general composition is $AE_{2-x}L_xSi_{1-y}O_{4-x-2y}N_x:RE$ where AE contains one or more elements selected from Mg, Ca, Sr, Ba and RE contains one or more elements selected from rare earth metals, preferably at least Eu, and L contains one or more elements selected from rare earth metals other than RE, with $0 < x \leq 0.1$, preferably $0.003 \leq x \leq 0.02$, and $0 < y \leq 0.1$, preferably $0.002 \leq y \leq 0.02$.

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It is also possible to use yellow/green-emitting beta-SiAlONs of the general formula $Si_{6-z}Al_zO_zN_{8-z}:Eu$ with $0 < z \leq 4$. The beta-SiAlONs may also have the general formula $Si_{6-x}Al_zO_yN_{8-y}:RE_z$ where $0 < x \leq 4$, $0 < y \leq 4$, $0 < z < 1$ and RE is one or more elements selected from rare earth metals, preferably at least Eu and/or Yb.

It is also possible to use yellow- to green-emitting nano-semiconductor materials, called "quantum dots", containing at least one compound selected from: a group II-VI compound, a group IV-VI compound or metal nanocrystals.

Inventive red-emitting phosphors may especially phosphors of the following general formula: $Sr(Sr_aCa_{1-a})Si_2Al_2N_6:D$ having an activator content of ≥ 4 mol%, preferably ≥ 8 mol%, further preferably ≥ 10 mol%, further preferably ≥ 15 mol%, where the divalent metals are preferably strontium and calcium with a calcium content of ≤ 15 mol%, preferably ≤ 10 mol%, further preferably ≤ 8 mol%, so as to result in the general formula $Sr(Sr_aCa_{1-a})Si_2Al_2N_6:D$ with $a \geq 0.7$, preferably $a \geq 0.8$, further preferably $a \geq 0.84$.

There follows a discussion of some embodiments of lighting devices discussed here in comparison with conventional lighting devices. The phosphors of inventive examples 1 and 2 shown in figures 48a to 48d were used in lighting devices having LEDs and are labeled with embodiment LED1 or embodiment LED2 in the figures which follow. The comparative examples were also labeled correspondingly. Figure 48a shows, in tabular form, spectral data of a comparative example 1, in which a conventional phosphor of the formula $CaAlSiN_3:Eu$ (0.4% Eu) has been incorporated in the standard silicone

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potting material of a blue-emitting LED. In contrast, in inventive examples 1 to 3, phosphors of the invention having different proportions of europium as activator were incorporated into the potting material. In the inventive examples, a smaller half-height width FWHM of the radiation emitted relative to comparative example 1 was observed, and inventive examples 1 and 2 simultaneously show a higher external quantum efficiency than the comparative example. The further figures 48b and 48c show the emission spectra of the comparative example and the three inventive examples 1 to 3. These emission spectra of the inventive examples show a reduced half-height width with simultaneously deep red emission.

Figure 48d shows the diffuse reflection of the comparative example and the three inventive examples as a function of wavelength. All the inventive examples comprising the phosphor of the invention exhibit a very low reflection in the UV to green region of the spectrum, which means a high absorption. At the same time, the reflection is very high at wavelengths > 650 nm, and so a high conversion efficiency can be achieved.

White light-emitting LED lighting devices were constructed with the aid of various combinations of embodiments of the present invention and comparative examples. A white point having the CIE coordinates CIE-x = 0.285 and CIE-y = 0.275 was chosen here. The resulting LED emission spectra were analyzed and compared, and the coverage of the color space was determined by employing a standard set of LCD filter absorption curves and determining the resulting filtered color points for the blue, green and red channels. It can be inferred from figure 48e that, in contrast to a comparative

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example 1, an inventive example 1 exhibits an elevated conversion efficiency and an increase in the luminous efficiency LER by 4%, with similar overlap with the sRGB color space in both lighting devices. Figure 48f shows the LED spectra of comparative example 1 and of inventive example 1 from the previous table in figure 48e. Figure 48g gives a comparison of the coverage of the sRGB color space for comparative example 1 and inventive example 1.

Figure 48h gives, in tabular form, the composition of a comparative example 2 and of inventive examples 2 and 3 for backlighting devices for an extremely large color space, for example DCI-P3. A white point with CIE-x = 0.275 and CIE-y = 0.250 was chosen. The resulting LED emission spectra were again analyzed and compared analogously to the examples already described above in figure 48e, but this time for the DCI-P3 color space. The overlap with this color space is comparable or higher in the case of the inventive examples. Figures 48i and 48j show the LED emission spectra of inventive examples 2 and 3 in comparison to comparative example 2 and the coverage of the DCI-P3 color space for these examples.

Second phosphors used may also be other phosphors, for example from the group of the SiONs, SiAlONs, silicates and quantum dots.

Embodiments of flash applications:

The phosphors of the invention, especially the phosphors of the general structural formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2(\text{N},\text{O})_6$, can also be used for flash applications together with the garnets activated by means of an activator, especially the above-

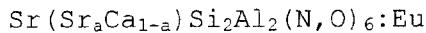
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described cerium-activated garnets. For this purpose, the radiation source used, which emits a primary radiation, is a blue-emitting LED, for example an InGaN LED, having a dominant wavelength of 300-500 nm, preferably 400-500 nm, further preferably 420-470 nm. A particularly suitable spectrum for a flash application, for example in mobile phone cameras, has an intensity, based on the maximum of the spectrum in the cyan color range (about 450-500 nm), of at least 12.5%. The intensity of the spectrum in the wavelength range of > 650 nm may at the same time be comparatively small, since typical sensors of cameras have a high sensitivity in this range and this spectral region is frequently filtered out by special IR filters in order to avoid disruptive influences of radiation from this radiation range on the sensor and the image quality.

For flash applications, it is possible here to use lighting devices having only one radiation source wherein the blue-emitting and/or UV radiation-emitting LED chip contains a phosphor mixture comprising at least one of the phosphors of the invention, for example together with a yellow/green-emitting garnet phosphor, in the beam path thereof. LED devices for flash applications may additionally at least also have two different LED modules, in which case one module emits a comparatively cold white light (correlated color temperature CCT between 4000-6000 K) and the further module a comparatively warm white light (correlated color temperature CCT between about 1500-3000 K). By controlled feeding of current to the two LED modules, even better variability of lighting situations of the environment in the subject of the photo is possible, for example in artificial light or in daylight.

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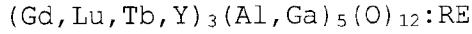
For a cold white color locus, the first phosphor used may be an inventive orange/red-emitting phosphor of the general formula:



with $a \geq 0.8$, preferably $a \geq 0.82$. The europium content is between 0.1-20 mol%, or between 1-10 mol%, further between 0.1-5 mol%, preferably between 0.1-3 mol%, further preferably between 0.1-2 mol%, based on the alkaline earth metals.

The emission peak of the phosphors of the invention may be between 600-640 nm, preferably between 605-625 nm, and the spectral half-height width at half the maximum height (FWHM) should be < 85 nm, preferably < 80 nm, additionally preferably < 78 nm. The emission intensity at wavelengths greater than 650 nm should be very low, since typical sensors of cameras have a high sensitivity in this range.

Second phosphors used for a cold white application may then be the above-described garnets of the general formula:



with RE = rare earth metals, especially Ce.

The garnets are especially blue/green- to yellow-emitting phosphors of the formulae $\text{Lu}_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{Ce}$ and $(\text{Lu},\text{Y})_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{Ce}$, which have particularly good excitability at a wavelength in the range of 425-455 nm, preferably 430-450 nm. Particular preference is given to a blue/green-emitting phosphor having very good stability and conversion efficiency at high temperatures and high radiation

intensities which are typical for flash applications, having the formula $\text{Lu}_3(\text{Al}_{1-x}\text{Ga}_x)_5(\text{O})_{12}:\text{Ce}$ with a cerium content of 0.5-5 mol%, preferably 0.5-2 mol%, based in each case on the rare earth metals, and a gallium content x of 0 to 0.5, preferably 0.15 to 0.3. Other garnets having other element combinations are likewise possible, especially variants in which some or all of the lutetium has been replaced by yttrium in the formula of the garnet $\text{Lu}_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{Ce}$. These combinations of a first and a second phosphor, compared to conventional combinations of phosphors where the above-described garnet phosphor has been mixed with another red-emitting phosphor from the class of the 2-5-8 phosphors of the general formula $(\text{Ca},\text{Sr},\text{Ba})_2(\text{Si},\text{Al})_5(\text{N},\text{O})_8:\text{Eu}$ with optional co-dopants, for example lanthanoids such as Mn, Nd, Dy, Sm, Tm and alkali metals such as Li, Na, K, have better stability in relation to the color point and higher LED efficiencies at elevated currents. Hereinafter, phosphors of the general formula $(\text{Ca},\text{Sr},\text{Ba})_2(\text{Si},\text{Al})_5(\text{N},\text{O})_8:\text{Eu}$ having optional co-dopants are referred to as "2-5-8 phosphors". In addition, phosphor mixtures of the invention exhibit reduced intensity of emission at wavelengths > 650 nm, but the intensity, based on the maximum of the spectrum in the cyan color region, of at least 12.5% is satisfied as an important condition for flash applications.

For warm white flash applications, it is preferable to use a yellow-emitting garnet phosphor $(\text{Gd},\text{Lu},\text{Tb},\text{Y})_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{RE}$, preferably of the formula $(\text{Gd},\text{Y})_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{Ce}$ or $(\text{Tb},\text{Y})_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{Ce}$, having maximum excitation in the range of 435-470 nm, preferably 440-465 nm. The preferred yellow-emitting phosphor has a very high stability and conversion efficiency at high temperatures and high radiation intensities (high currents) which are typical of flash

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applications. A particularly preferred yellow/green-emitting phosphor is $Y_3(Al_{1-x}Ga_x)_5(O)_{12}$:Ce with a cerium content of 1.5-5 mol%, preferably 2.5-5 mol%, and a gallium content x of 0 to 0.5, preferably x of 0 to 0.1. Other element combinations within the $(Gd, Lu, Tb, Y)_3(Al, Ga)_5(O)_{12}$:Ce system having similar spectral characteristics are likewise possible.

LED lighting devices suitable for flash applications may, independently of lighting devices having just one LED, for example, at least also have two or three blue-emitting LED chips as primary radiation-emitting radiation sources, with the phosphors and phosphor mixtures already mentioned above having been incorporated in the beam path thereof. The CIE color gamut of the converted light of the LED lighting device is preferably on the line of a blackbody emitter (Planck) in the range from 6500 K to 2700 K with a deviation of 3 steps of a MacAdam ellipse, more preferably in the range from 5000 K to 3000 K.

In the case of an LED lighting device having two LED chips as radiation sources, in a further embodiment, the radiation from the first LED chips, after conversion, has a CIE color locus within a range enclosed by the following CIE color coordinates (Cx/Cy): (0.21; 0.247), (0.26; 0.24), (0.24; 0.32), (0.28; 0.31). The second LED chip as the second radiation source, after conversion, has a CIE color locus which is enclosed by the following CIE coordinates: (0.45; 0.41), (0.46; 0.39), (0.58; 0.39), and (0.58; 0.42). In such an LED lighting device, the individual radiation sources can be operated with different driver currents, advantageously with mixing of the converted light emitted in an optical element, such as a common lens, to give an overall emission radiation.

In the case of an LED lighting device which is suitable for flash applications and has three LED chips as radiation sources, in a further embodiment, the first two LED chips have the CIE color loci already described above and the third radiation source, the third LED module, after conversion, has a CIE color locus which is enclosed by the following coordinates: (0.40; 0,44), (0.39; 0.51), (0.45; 0.52), and (0.47; 0.46). In this embodiment too, the converted light emitted is mixed by a downstream optical element in the beam path, such as a lens.

The phosphor particles preferably have a median particle size d_{50} of 5-30 μm , more preferably 7-17 μm . The particle size distribution can be determined, for example, via laser diffraction by means of the Fraunhofer approximation which is known to those skilled in the art.

Some inventive examples of LED lighting devices suitable for flash applications are to be elucidated in detail hereinafter. In a comparative example 1, an inventive example 1 and a comparative example 2 and an inventive example 2, different phosphor mixtures according to the prior art and phosphor mixtures comprising phosphors of the invention are incorporated in each case into a standard silicone potting material of a blue-emitting InGaN-LED chip as radiation source. The illumination area of each of the LED chips is 1 mm^2 .

In comparative example 1, 11.5 percent by weight of phosphor based on the silicone potting material is incorporated, using $\text{Lu}_3\text{Al}_4\text{GaO}_{12}:\text{Ce}$ having a cerium content of 1.5 mol% based on the rare earth metals as green-emitting phosphor. The red-

emitting phosphor used is a conventional 2-5-8 phosphor $\text{SrBaSi}_5\text{N}_8:\text{Eu}$ having an Eu content of 1.5 mol% based on the alkaline earth metals. The dominant wavelength of the LED chip at 350 mA is 447 nm and the ratio of the green/red phosphors is 5.7:1. As inventive example 1, the same green-emitting phosphor as used in comparative example 1 is used, and the red phosphor used is an inventive phosphor $\text{Sr}(\text{Sr}_{0.86}\text{Ca}_{0.14})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ having an Eu content of 0.4 mol%, based on the rare earth metals. The ratio of the green/red phosphors is 3.9:1. The dominant wavelength of the LED chip at 350 mA is again 447 nm.

The table in figure 49a shows the respective x and y CIE color coordinates of comparative example 1 and of inventive example 1 at different currents (average of 4 LEDs), with comparative example 1 set at the percentage of 100% for the corresponding currents. It can be inferred from this table, but in particular also from figures 49b and 49c, that, with increasing current, the phosphor mixtures comprising the phosphors of the invention are more stable than the conventional phosphor mixtures, the conventional phosphor mixtures losing some of their relative red emission intensity compared to the green-yellow emission intensity, whereas the phosphor mixtures of the invention remain virtually stable. It can additionally be inferred from the table that the LED comprising the phosphor of the invention has a higher conversion efficiency as the ratio of the luminous flux Φ_v of an LED having a potting material filled with the phosphor mixtures and the radiant power Φ_e of an LED having a clear potting material without phosphors $\Phi_{(v(\text{filled potting}))}/\Phi_{e(\text{clear potting})}$ compared to the LED comprising the conventional phosphor mixture. The luminous efficiency LER is defined as:

$$LER = \frac{\int_{380\text{nm}}^{780\text{nm}} \Phi_{e(\text{filled potting})}(\lambda) V(\lambda) d\lambda}{\int_0^{\infty} \Phi_{e(\text{filled potting})}(\lambda) d\lambda}$$

LED spectra of comparative example 1 and of inventive example 1 at currents of 40 and 1000 mA are shown in figures 49b and 49c. It is clearly apparent here that the conventional phosphor mixture, with increasing current, loses emission in the red wavelength range (fig. 49b), whereas there is only a very slight decrease in the case of the phosphor mixture of the invention (fig. 49c).

The effect observed can be attributed to the different red-emitting phosphors in the conventional and inventive phosphor mixtures. In the case of the 2-5-8 phosphors used in the conventional phosphor mixture, a decrease in the conversion efficiency with increasing current is observed, which is manifested in a decrease in the red emission in relation to the yellow/green emission in the LED spectrum in conventional phosphor mixtures with increasing current. In the case of phosphor mixtures comprising the phosphor of the invention, a distinctly reduced decrease by comparison in the conversion efficiency of the red phosphor with increasing current is observed. The relative decrease in the conversion efficiency with increasing current for a phosphor of the invention, such as for a conventional 2-5-8 phosphor, is shown in figure 49d.

Figure 49e shows a comparison of the LED spectra normalized to the maximum emission intensity for inventive example 1 and for comparative example 1. Inventive example 1 shows a reduced emission intensity within a wavelength range of > 650 nm and simultaneously has a relative emission intensity of > 12.5% in the cyan region.

Figure 49f shows the normalized emission intensity of a typical phosphor, by way of example a 2-5-8 phosphor, and of

a phosphor of the invention. What is clearly apparent here is the reduced emission intensity of the phosphor of the invention in the wavelength range of > 650 nm, which is attributable to the reduced FWHM.

The shift in the color point of the LEDs of comparative example 1 and of inventive example 1 with increasing current is shown in figure 49g. There is a much smaller shift here in the color point of the LED of inventive example 1 compared to the LED of comparative example 1 because of the higher emission intensity of the phosphor of the invention.

In a further comparison of an inventive example 2 compared with a conventional comparative example 2, a phosphor mixture having an amber color point is used. Comparative example 2 comprises a cerium-activated garnet phosphor $(Y_{0.957}Ce_{0.043})_3Al_5O_{12}$ together with a conventional 2-5-8 phosphor $(Ca_{0.1}Sr_{0.4}Ba_{0.5})_2Si_5N_8:Eu$ having an Eu content of 3.25 mol% based on the alkaline earth metals used. The concentration of the phosphors is 41% by weight based on the silicone potting material and the dominant wavelength of the blue LED chip at a current of 350 mA is 444.7 nm. The ratio of the yellow/red phosphors is 5.9:1 (% by weight ratios). In inventive example 2, the same garnet phosphor as in comparative example 2 is used, but in that case an $Sr(Sr_{0.86}Ca_{0.14})Si_2Al_2N_6:Eu$ phosphor of the invention having an Eu content of 2 mol% based on the rare earth metals is used in a yellow/red ratio of 5:1. The dominant wavelength of the blue-emitting LED chip at a current of 350 mA is 444.5 nm, with use of 39% by weight of phosphor mixture in relation to the silicone potting material.

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It can be inferred from the table in figure 50a that the LED comprising the phosphor of the invention again has a higher conversion efficiency $\Phi_v(\text{filled potting}) / \Phi_e(\text{clear potting})$ compared to the LED comprising the conventional phosphor mixture at the respective currents. These values are reported in relation to comparative example 2 that has been normalized to 100% for the respective current (average of 4 LEDs).

In the most recent generation of mobile communications devices, it is possible to use "true-tone flashes" which, as well as a cold white light-emitting LED, also have a second LED which emits either warm white light or yellow light ("amber"). This second LED is used in situations where the ambient light has a lower color temperature than daylight (cold white), in order to produce a flash which approximates as closely as possible to the color temperature of the ambient light. In order to produce light of various color temperatures, the cold white light and the light of the yellow-emitting LED have to be mixed with one another in various ratios. This can be achieved, for example, by operating the two LEDs with different currents. When the flash has a similar color temperature to the ambient light, the colors in the image have a more natural appearance.

Similarly to inventive example 1, the inventive phosphor mixtures of inventive example 2 are also more stable with increasing current than the conventional phosphor mixtures, which lose a considerable proportion of their red emission compared to the yellow emission components, particularly at relatively high currents.

Figures 50b and 50c show the LED spectra of comparative example 2 (figure 50b) and of inventive example 2 (figure

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50c) at currents of 40 and 1000 mA. A comparison of the two LED spectra shows that, in the case of the conventional phosphor mixture of comparative example 2, a considerable proportion of the red-emitting spectrum is lost with rising currents compared to phosphor mixtures of the invention.

Figure 50d shows the stability of the color point with increasing currents for comparative example 2 and inventive example 2. Because of the smaller decrease in emission in the red region in inventive example 2 in relation to comparative example 2, the change in the color point of the LED of inventive example 2 is much less marked than in the case of the LED of comparative example 2.

The normalized LED spectra for inventive example 2 and comparative example 2 are shown in figure 50d. The phosphor mixture of the invention, compared to the conventional phosphor mixture, shows a significant reduction in emission intensity in the wavelength range of > 650 nm. This is attributable particularly to the reduced FWHM of the phosphor of the invention compared to the conventional 2-5-8 phosphor.

In summary, it can be stated that it is therefore possible to achieve an equal light yield with lower operating currents in the case of phosphor mixtures of the invention compared to conventional phosphor mixtures. Since the current consumption in mobile applications, for example mobile phones, is a crucial criterion for operating life, a low consumption is very important. In addition, brighter LEDs are possible with the phosphor mixtures of the invention, which extends the range of color points, especially for "true-tone flashes". Should a higher light yield not be desirable, it is possible to use LED chips with weaker emission of radiation compared

to conventional LED chips, for example for yellow-emitting LEDs ("amber"), which reduces rejects during LED chip production and hence also enables the utilization of LED chips that are otherwise too dark.

Embodiments of warm white light with CRI \geq 80:

In a further embodiment of the present invention, phosphors of the invention are used for production of warm white light, for example for general lighting applications. Warm white light-emitting lighting devices comprising phosphors of the invention can especially achieve a "color rendering index" (CRI) of ≥ 80 , preferably ≥ 82 .

Radiation sources of particularly good usability for lighting devices for production of warm white light may be blue-emitting LEDs, for example InGaN-LEDs, which emit light having a dominant emission wavelength of 430-470 nm, preferably 440-460 nm, further preferably of 445-455 nm, as primary radiation. First phosphors used for conversion of the primary radiation may be inventive phosphors of the general formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{D}$ where $0.7 \leq a$, preferably $0.8 \leq a$, further preferably $0.84 \leq a$, where $a < 1$ and where the proportion of activator D, preferably europium, based on the molar proportion of the alkaline earth metals, is ≥ 1 mol%, preferably ≥ 2 mol%, further preferably ≥ 3 mol%.

Second phosphors used may, for example, be a green/yellow-emitting garnet of the general formula $(\text{Gd},\text{Lu},\text{Y},\text{Tb})_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{RE}$ with RE = rare earth metal, preferably Ce. The garnet preferably has the general formula $\text{Y}_3(\text{Al}_{1-x}\text{Ga}_x)_5(\text{O})_{12}:\text{Ce}$ where the proportion of Ga is $0.2 \leq x \leq$

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0.6, preferably $0.3 \leq x \leq 0.5$, further preferably $0.35 \leq x \leq 0.45$.

One advantage of using a garnet phosphor of the formula $Y_3(Al_{1-x}Ga_x)_5(O)_{12}:Ce$ rather than a garnet phosphor of the formula $Lu_3(Al_{1-x}Ga_x)_5(O)_{12}:Ce$ is that the first garnet phosphor $Y_3(Al_{1-x}Ga_x)_5(O)_{12}:Ce$ has a lower density of about $4.5-5 \text{ g/cm}^3$, while the second phosphor $Lu_3(Al_{1-x}Ga_x)_5(O)_{12}:Ce$ has a density of about $6.7-7 \text{ g/cm}^3$, and therefore a smaller mass of phosphor is consumed for a given application. In addition, therefore, the cheaper and more widely available Y_2O_3 can be used rather than Lu_2O_3 as starting material for the production of the phosphor, such that the procurement cost for the garnet phosphor is reduced.

The use of a first phosphor of the invention exhibits higher absorptions compared to phosphor mixtures containing conventional 2-5-8 phosphors or $CaAlSiN_3$ phosphors.

Surprisingly, the absorption in the case of phosphors of the invention, given the same activator content, is considerably higher than in the case of conventional 2-5-8 phosphors. This enables a drastic reduction in the amount of red-emitting phosphor compared to conventional solutions and a very high conversion efficiency. At the same time, the excellent optical properties of the red-emitting phosphors of the invention enable a very high light yield and high conversion efficiency with high CRI.

Some embodiments of lighting devices of the invention for general lighting applications having a high CRI are to be described in detail hereinafter.

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Figure 51a shows a tabular compilation of comparative examples 1 and 2 and of inventive examples 1 and 2. In all examples, a blue-emitting InGaN-LED having a dominant wavelength of 446 nm is used, wherein the standard silicone potting material incorporates phosphor mixtures in which garnet phosphors have been mixed either with conventional red-emitting phosphors or with red-emitting phosphors of the invention in the silicone potting material.

In comparative example 1, a cerium-activated yttrium aluminum gallium garnet (abbreviated to YAGaG in the figures which follow) with a conventional $(\text{Sr}_{0.7}\text{Ca}_{0.3})\text{AlSiN}_3$ phosphor (abbreviated to SCASN in the figures which follow) is incorporated as phosphor mixture in a concentration of 15% by weight relative to the total mass of the silicone potting material of the blue LED, where the ratio of green to red phosphor is 2.0 (% by weight ratio). The absolute concentration of green phosphor is 10% by weight and that of red phosphor 5% by weight. In addition, the correlated color temperature CCT, the CRI, the R9 value for red hues and the relative conversion efficiency relative to comparative example 2 are reported.

Comparative example 2 contains a mixture of a lutetium aluminum garnet (abbreviated to LuAGaG in the figures which follow) and a conventional 2-5-8 phosphor (abbreviated to 258 in the figures which follow). In contrast, inventive examples 1 and 2 contain either an yttrium aluminum garnet or a lutetium aluminum garnet together with different phosphors of the invention (abbreviated to 226 in the figures which follow).

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It can be inferred from the tabular listing in figure 51a that all comparative and inventive examples have a correlated color temperature CCT in the range of $2700\text{ K} \pm 15\text{ K}$ with a high CRI ≥ 80 and a high R9 of 10 ± 1 . Inventive examples 1 and 2 show an elevated conversion efficiency in relation to comparative examples 1 and 2, but smaller amounts of red-emitting phosphors are required than in the comparative examples. Particular preference is given to inventive example 1, since least red-emitting phosphor is used therein and, in addition, an yttrium aluminum gallium garnet is also used, which avoids the above-described high costs of the lutetium aluminum garnet.

Figure 51b shows an amount in %, based on comparative example 1, of red-emitting phosphor which has to be expended for the inventive and comparative examples. It is clearly apparent here that, for inventive examples 1 and 2 of the present invention, much less red-emitting phosphor has to be used than for the comparative examples.

The phosphors of the invention exhibit a very low spectral half-height width at half the maximum height FWHM compared to the conventional red-emitting phosphors, as apparent from figure 51c.

Figure 51d shows the emission spectra of the green-emitting garnet phosphors of the present comparative and inventive examples. It is apparent here that the yttrium aluminum garnet exhibits an emission intensity in the blue/green region of the visible spectrum (470-520 nm) which is comparable to or better than with other green-emitting garnet phosphors. For this reason, it is possible with this garnet phosphor, in a particularly inexpensive manner (avoidance of

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lutetium), to implement warm white-emitting lighting devices having a high CRI.

The thermal quenching of various green/yellow-emitting garnet phosphors and a green-emitting orthosilicate phosphor from room temperature to 125°C in relation to the absolute brightness at 25°C is shown in figure 51e. It can be inferred from this diagram that the preferred phosphor of the $(\text{Lu},\text{Y})_3(\text{Al},\text{Ga})_5(\text{O})_{12}:\text{Ce}^{3+}$ type, in contrast to orthosilicate phosphors, exhibits only very low thermal quenching at relatively high temperatures.

The adverse effects of the thermal quenching of the various red-emitting phosphors used in the comparative and inventive examples on the absolute brightness are shown in figure 51f. In this figure, the phosphors of the invention show comparable thermal quenching to the best 2-5-8-phosphors, while another 2-5-8 phosphor $\text{Ca}_2\text{Si}_5\text{N}_8:\text{Eu}^{2+}$ (2%) exhibits appreciable quenching.

Figures 51g and 51h show the LED spectra of the LEDs of inventive examples 1 and 2. In these spectra, the peaks of the unconverted blue primary radiation from the LED at wavelengths between 410 and 460 nm, and the green-red components of the secondary radiation converted are clearly apparent. The additive color mixing of these primary and secondary radiation components produces warm white light with a high CRI.

In a further embodiment of the present invention, a lighting device for production of a white light having a $\text{CRI} \geq 90$ is provided, wherein the radiation source emits a primary radiation in the wavelength range between 430 nm and 470 nm,

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preferably 440 and 460 nm, further preferably 445 nm and 455 nm, and the second phosphor present is a garnet of the general formula $(\text{Gd}, \text{Lu}, \text{Y}, \text{Tb})_3(\text{Al}, \text{Ga})_5(\text{O})_{12}:\text{RE}$, preferably $(\text{Lu}, \text{Y})_3(\text{Al}, \text{Ga})_5(\text{O})_{12}:\text{RE}$ with RE = rare earth metal, preferably Ce.

In this lighting device for production of a white light having a $\text{CRI} \geq 90$, in the first phosphor, which may have one of the general formulae already described above, especially $\text{Sr}(\text{Sr}_a\text{M}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{D}$, the metal M is Sr and Ca, where the parameter a is as follows: $0.7 \leq a$, preferably $0.8 \leq a$, further preferably $0.84 \leq a$, where the proportion of the activator D is $\geq 1.5\%$, preferably $\geq 3.5\%$, further preferably $\geq 4.5\% \text{ mol\%}$.

Some embodiments of phosphor mixtures of the invention having a high $\text{CRI} \geq 90$ are to be presented hereinafter.

Figure 52a shows, in tabular overview, a comparative example 1 and inventive examples 1 to 4, which, as well as a cerium-activated yellow/green-emitting garnet phosphor, also contain phosphors of the invention. The primary radiation source used was an InGaN-LED chip having the dominant wavelengths specified, with the phosphor mixtures disposed in the beam path thereof (phosphor mixtures present in the standard silicone potting material). The measurements were conducted at a correlated color temperature (CCT) of $2700 \text{ K} \pm 30 \text{ K}$. It is apparent here that inventive examples 1 to 4 have an elevated conversion efficiency compared to the conventional comparative example. At the same time, less red phosphor has to be used in the phosphor mixtures of the invention.

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With reference to the preceding figure 52a, figure 52b shows a comparison of the measurement data for comparative example 1 as opposed to inventive example 1 at correlated color temperatures of 2700 K and 4000 K. It is again apparent that the conversion efficiency of the inventive example is higher than the conversion efficiency of the comparative example.

Figure 52c shows, in the left half, a graph of the reflectivity as a function of the wavelength for a conventional phosphor $(\text{Sr,Ca})\text{AlSiN}_3$: 0.4% Eu, and for the two phosphors of the invention $\text{Sr}(\text{Sr}_{0.84}\text{Ca}_{0.16})\text{Si}_2\text{Al}_2\text{N}_6$: 4.7% Eu (labeled as $\text{Sr}(\text{Sr,Ca})\text{Si}_2\text{Al}_2\text{N}_6$:Eu (8% Ca, 4.7% Eu) in the diagram) and $\text{Sr}(\text{Sr}_{0.8}\text{Ca}_{0.20})\text{Si}_2\text{Al}_2\text{N}_6$: 3.7% Eu (labeled as $\text{Sr}(\text{Sr,Ca})\text{Si}_2\text{Al}_2\text{N}_6$:Eu (10% Ca, 3.7% Eu) in the diagram). What is noticeable here is the higher absorption of the phosphors of the invention in the wavelength range between 300-600 nm because of the elevated content of europium. At the same time, the phosphors of the invention exhibit an elevated conversion efficiency. The right-hand image in figure 52c shows the high consumption of red phosphor in comparative example 1 (left-hand bar) as compared with inventive example 1 (right-hand bar) at correlated color temperatures of 4000 and 2700 K, and it becomes particularly clear here that, in the case of phosphor mixtures of the invention containing the novel phosphor, significantly less phosphor has to be used.

A graph comparison of the temperature-dependent changes in the LED color locus of two embodiments of the present invention compared to comparative example 1 from room temperature to 85°C is shown in figure 52d. It is apparent here that the shift in the LED color locus as a function of temperature is much more marked in comparative example 1 than in inventive examples 3 and 4 of the present invention.

Figure 52e shows a comparison of the emission spectra of two red-emitting phosphors of the invention compared to a conventional phosphor of the formula $(\text{Sr,Ca})\text{AlSiN}_3$: 0.4% Eu. The phosphors of the invention, by contrast with the conventional phosphor, show lower half-height widths FWHM, which result in a high color rendering index (CRI) and in an increased efficiency.

The emission spectra of various green-emitting garnet phosphors at an excitation wavelength of 460 nm, which are used in phosphor mixtures of the invention for warm white light applications, are shown in figure 52f. The emission intensity of these garnet phosphors in the blue/green region of the visible spectrum (470-520 nm) is either comparable to or even better compared to other green-emitting garnet phosphors. This allows a good color rendering index (high CRI).

Figures 52g and 52h show spectra of blue-emitting LEDs with phosphor mixtures according to inventive example 1 having been introduced into potting material thereof at correlated color temperatures of 2700 K (figure 52g) and at 4000 K (figure 52h). In both spectra, the signals of the secondary radiation of the phosphors of the invention in the red and green region and also the emission of the unconverted primary radiation of the LED in the blue region are clearly apparent. Further embodiments of the present invention are directed to phosphor mixtures or lighting devices in which at least three phosphors are disposed in the beam path of the radiation source, for example of a blue LED. In order to adjust either the CRI or the LED efficiency for a given color locus, it is possible to use phosphor mixtures having more than two

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phosphors. Especially in the case of use of three phosphors, for example of a green-emitting phosphor, a yellow-emitting phosphor and a red-emitting phosphor, there are several ways of obtaining an LED having a particular color point. However, one problem in the prior art is that many conventional orange/red-emitting phosphors have a broad-band emission and a considerable portion of the red light is emitted within a range to which the human eye is relatively insensitive.

What are proposed, therefore, are phosphor mixtures including at least one phosphor of the invention. Such lighting devices of the invention therefore have, as radiation source for the primary radiation, a blue LED having a dominant wavelength of 300-500 nm, preferably 400-500 nm, further preferably 420-470 nm. The phosphor mixture contains at least one red-emitting phosphor of the invention as the first phosphor having the general formula $Sr(Sr_aM_{1-a})Si_2Al_2N_6:D$ where M is selected from the group of Ca, Ba, Zn, Mg, preferably Ca, a yellow/green-emitting garnet phosphor of the general formula $(Y, Lu, Gd, Tb)_3(Al, Ga)_5O_{12}:Ce$ having a peak wavelength of 500-570 nm, preferably of 510-560 nm, further preferably of 520-550 nm, as the second phosphor, and, as the third phosphor, either an orange/red-emitting inventive phosphor of the general formula $Sr(Sr_aM_{1-a})Si_2Al_2N_6:D$ where M is selected from the group of Ca, Ba, Zn, Mg, preferably Ca, or an orange/red-emitting 2-5-8 phosphor of the general formula $M_2(Si, Al)_5(N, O)_8:Eu$ with M = Ca, Sr, Ba or a yellow-emitting garnet phosphor of the general formula $(Y, Lu, Gd, Tb)_3(Al, Ga)_5O_{12}:Ce$ having a peak emission wavelength of 580-650 nm, preferably of 590-640 nm, further preferably of 600-625 nm for the phosphor of the invention or the 2-5-8 phosphor, and of 500 to 600 nm, preferably 525 to 575 nm, further preferably of 535 to 565 nm, for the garnet phosphor.

The red-emitting phosphor of the invention as the first phosphor is preferably selected such that it gives, in combination with the blue LED and the garnet phosphor and the yellow/red phosphor, a CRI of ≥ 75 , preferably ≥ 80 , further preferably ≥ 85 and more preferably ≥ 90 for a wide range of correlated color temperature CCT of at least 4000 to 2700 K, further preferably of 5000 to 2700 K, additionally preferably of 6500 to 2400 K. This is best achieved by using a phosphor of the invention having a half-height width FWHM of < 85 nm, preferably < 82 nm, further preferably < 80 nm, having the general formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a \geq 0.8$ and a europium content of 0.1 to 10 mol%, preferably 2-5 mol%. The preferred phosphors of the invention have a high thermal stability and a high conversion efficiency under typical operating conditions for warm white lighting devices.

The green/yellow-emitting garnet phosphor used as the second phosphor may, for example, have the general formula $\text{Lu}_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$, and it is selected so as to result in a high conversion efficiency and thermal stability. This can be done, for example, by selecting a garnet phosphor of the general formula $\text{Lu}_3(\text{Al},\text{Ga})_5\text{O}_{12}:\text{Ce}$ having a cerium content of 1-5 mol%, preferably 1-3 mol%, and a gallium content of 0-50 mol%, preferably 0-30 mol%, so as to result in the general formula $\text{Lu}_3(\text{Al}_{1-x}\text{Ga}_x)_5\text{O}_{12}:\text{Ce}$ with $0 \leq x \leq 0.5$, preferably $0 \leq x \leq 0.3$. Other element combinations within the general system of the garnet phosphors are likewise possible, especially variants in which at least some of the lutetium is replaced by yttrium in the general formula.

The third phosphor used may preferably be a phosphor which, in combination with the blue LED and the garnet phosphor, and

also the red-emitting phosphor of the invention, gives a CRI of ≥ 75 , preferably ≥ 80 , further preferably ≥ 85 and most preferably ≥ 90 over a wide CCT range, for example of 4000-2700 K, further preferably of 5000-2700 K, most preferably of 6500-2400 K. For example, the third phosphor used may be a nitridosilicate phosphor of the general formula $(\text{Ca}, \text{Sr}, \text{Ba})_2(\text{Si}, \text{Al})_5(\text{N}, \text{O})_8:\text{Eu}$, where the peak wavelength is 580-650 nm, preferably 590-640 nm, further preferably 600-625 nm, with a europium content of 0.1 to 10 mol%, preferably of 0.1 to 5 mol%, additionally preferably of 0.5-3 mol%, and a barium content of 30-100 mol%, preferably of 40-75 mol%, further preferably of 45-55 mol%, and a calcium content of 0-20 mol%, preferably 0-10 mol%, additionally preferably 0-5 mol%, where the strontium content is selected such that it adds up to 100 together with the alkaline earth metals and the europium.

Alternatively, the third phosphor used may also be an inventive phosphor of the general formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ having a half-height width FWHM of < 85 nm, preferably < 80 nm, additionally preferably < 78 nm, where the peak wavelength is 580-650 nm, preferably 590-640 nm, further preferably 600-625 nm. In order to achieve these spectral properties, an inventive phosphor of the general formula $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ with $a \geq 0.8$, preferably $a \geq 0.82$, further preferably $a \geq 0.85$, with a europium content of 0.1 to 5 mol%, preferably 0.1 to 3 mol%, most preferably 0.1 to 2 mol%, is used.

The third phosphor used may additionally be a yellow-emitting garnet phosphor $(\text{Lu}, \text{Gd}, \text{Tb}, \text{Y})_3(\text{Al}, \text{Ga})_5\text{O}_{12}:\text{Ce}$ having a peak emission wavelength of 500-600 nm, preferably 525-575 nm, further preferably of 535-565 nm. This can especially be

accomplished by a garnet phosphor of the general formula $Y_3(Al, Ga)_5O_{12}:Ce$ having a high conversion efficiency and thermal stability. This can be accomplished, for example, by using a garnet phosphor of the general formula having a cerium content of 1 to 6 mol%, preferably 1 to 4 mol%, and a gallium content of 0-50 mol%, preferably 0-25 mol%, so as to result in the general formula $Y_3(Al_{1-x}Ga_x)_5O_{12}:Ce$ with $0 \leq x \leq 0.5$, preferably $0 \leq x \leq 0.25$, and other element combinations are additionally also possible within this system, for example at least partial replacement of yttrium by lutetium.

Particular technical advantages of such lighting devices of the invention according to inventive examples 1 and 2 compared to comparative examples 1 and 2 are to be elucidated hereinafter.

Figure 53a shows the composition and the concentrations of the phosphor mixtures and particular ratios of the first to third phosphors in lighting devices having a standard InGaN LED with a chip area of 1 mm². Figure 53b shows the respective CIE color coordinates and the CRI and the corresponding conversion efficiencies, these having been expressed in relation to comparative example 1 which has been set to 100%. It is again clearly apparent here that inventive examples 1 and 2 have higher LED efficiencies than the conventional comparative examples.

Figure 53c shows a comparison of the LED spectra of comparative examples 1 and 2 and of inventive examples 1 and 2 together with the eye sensitivity curve for daytime vision of the human eye. It is apparent here that a high proportion of the rise in the LED efficiency is connected to the rise in the luminous efficiency which arises from the use of

phosphors of the invention. More particularly, the emission intensity is reduced in a spectral region in which the human eye has barely any sensitivity to light by use of the phosphors of the invention. Figure 53d shows the relative LED brightness for the comparative and inventive examples as a function of temperature. The lighting devices of the invention have a lower loss of brightness at high temperatures compared to the brightness at 23°C than the comparative examples.

Embodiments of LARP applications:

In a further embodiment of the present invention, the radiation source used which emits the primary light may also be a laser, for example a laser diode. In this case, it is advantageous when the first phosphor of the invention is spaced apart from the laser radiation source (LARP; "laser activated remote phosphor"). LARP applications of this kind are known, for example, from PCT patent applications WO 2012/076296 A2, WO 2011/098164 A1 and WO 2013/110495 A2, and the further patent applications DE 10 2012 209 172 A1, DE 10 2010 062 465 A1, DE 10 2011 002 961 A1 and DE 10 2012 201 790 A1. With lighting devices of this kind, for example projectors, it is possible to achieve significantly higher luminances than with conventional radiation sources.

In the embodiments which follow, a blue LED laser diode having a radiation density of 8.9 W/mm² and a peak wavelength of 446 nm is used, the beam of which is directed onto either conventional or inventive phosphor grains present at the base of an Ulbricht sphere having a reflective inner surface. The converted light reflected is subsequently collected and

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analyzed. Such a test setup simulates the conditions in an LARP lighting device.

In the table in figure 54, in a comparative example 6, a conventional phosphor is irradiated with the laser light, while, in inventive example 7, an embodiment of a phosphor of the invention is irradiated with laser light. It is clearly apparent that, given similar dominant wavelengths of 601 nm and 597 nm of the converted light, the luminous efficiency in the case of the phosphor of the invention is 42% higher than in the case of the conventional phosphor. The corresponding emission spectrum is shown in figure 55, the dotted line corresponding to comparative example 6 and the solid line to inventive example 7.

In the table in figure 56 and the corresponding emission spectrum in figure 57 too, a conventional phosphor and the same phosphor of the invention were irradiated with laser light in a comparative example 7 and an inventive example 8, as in the table in figure 54. In these experiments too, again, an elevated luminous efficiency is found with the phosphor of the invention compared to the conventional phosphor (increase by 13%).

More particularly, in all flash applications and also in other conversion applications of the phosphors of the invention, for example general lighting applications, and provided that the applications are not full-conversion applications, it is also possible for proportions of unconverted primary radiation from the radiation source to be present, even if this is not mentioned explicitly. Mixing of this unconverted primary radiation with the converted secondary radiation results in an overall emission of the

lighting device. As already described further up, for example, it is possible to achieve warm white light-emitting lighting applications with phosphors of the invention by converting the blue primary radiation from InGaN LEDs by means of the phosphors of the invention to a red component, and by means of green/yellow-emitting phosphors to a green-yellow component, in which case additive color mixing of the blue primary radiation and the two converted secondary radiations gives warm white light as the overall emission of the lighting device.

Embodiments with different Eu dopant concentrations:

Figures 58a to 58c show the dominant wavelengths, the relative intensities of photoluminescence, and the relative photoluminescence intensities assessed by eye of various embodiments of an inventive phosphor of the formula $\text{Sr}(\text{Sr}_{0.86}\text{Ca}_{0.14})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ as a function of rising activator concentrations of europium. With rising concentration of europium dopant, the dominant wavelength of the emission of the phosphors of the invention moves to higher wavelengths, from the orange to the red color range (see figure 58a), with a rise in the relative intensity of photoluminescence of 0.1 to about 4 mol% and then a drop again as the activator concentrations of europium increase further (see figure 58b). Based on the relative intensity of photoluminescence, a concentration range of 1-10 mol% of Eu, or 2-5 mol%, is preferred. Substantially analogous behavior to the relative intensity of the photoluminescence is also displayed by the relative luminescent photoluminescence intensity assessed by eye sensitivity, which likewise increases with rising activator concentrations of europium and drops again from about 4 mol% to about 20 mol% (see figure 58c). This takes

account of the sensitivity of the human eye for daytime vision. Based on the luminescent photoluminescence intensity, activator concentrations of 0.4-10 mol% of Eu, or 1-5 mol% of europium, are preferred.

Embodiments with further co-dopants or dopants other than Eu:

Further embodiments of phosphors of the invention having the general formulae

$\text{Sr}_{(1-x-h)}(\text{Sr}_a\text{M}_{1-a})_{(1-y-i)}\text{A}_{(x+y)}\text{B}_{(h+i)/2}\text{E}_{(h+i)/2}\text{Si}_{(2-z)}\text{G}_z\text{Al}_2\text{N}_6:\text{D}$ and

$\text{Sr}_{(1-x)}(\text{Sr}_a\text{M}_{1-a})_{(1-y)}\text{B}_{(x+y)}\text{Si}_{2-(x+y)}\text{Al}_{2+(x+y)}\text{N}_6:\text{D}$ are to be described in detail hereinafter.

Figure 59a shows the nominal composition of five different embodiments of a phosphor of the invention, the first phosphor having Cu co-doping and the second phosphor having Zn co-doping, and with replacement of the alkaline earth metals Sr and Ca by a mixture of La and Li in the third and fourth phosphors. In the last phosphor, La and Al replace the alkaline earth metals Sr and Ca, and also Si. The table reports the spectral properties of the various phosphors, especially the color locus in the CIE color space, the average reflectance between 450 and 470 nm (R(450-470)), the luminous efficiency LE (LE = LER/683 [lm/W] where LER = luminous efficacy), the dominant wavelength λ_{dom} and the spectral width at half the maximum height FWHM. It is clearly apparent how the co-doping with different metals affects the luminous efficiency and also all other spectral properties of these phosphors. Figures 59b to 59f which follow show the emission spectra of these phosphors of the invention.

Figure 59g shows a collation of the x-ray diffractograms of the above-described co-doped phosphors of the invention. An arrow in each case marks the characteristic x-ray reflection

which is one of the factors responsible for the assignment of the crystal structure of these co-doped phosphors of the invention too into the monoclinic $P2_1$ space group.

Figure 60a shows, in tabular overview, various embodiments of phosphors of the invention in which no carbon is present and, in the two other cases, different amounts of carbon are present for co-doping. The phosphors each have the same activator concentration of 3 mol% of Eu based on the alkaline earth metals. Analogously to figure 59a, the spectral properties of the various phosphors are listed here again too, "centroid WL" referring to the centroid wavelength of the emission spectrum, which is a weighted average of the frequencies present in the emission spectrum. The table shows that co-doping with carbon causes a red shift in the emission spectrum of the phosphors, which can be used, for example, to improve the color rendering index of lighting devices of the invention. Figure 60b which follows shows the emission spectra of the various phosphors in the table in figure 60a. What is clearly apparent here is the red shift on account of the co-doping with carbon.

The table in figure 61a shows various embodiments of phosphors of the invention with different activators. In the first phosphor in the table, a mixture of Eu, Ce and Li is used, while either manganese Mn or cerium Ce serves as activator in the case of the other phosphors. It is clearly apparent that the different activators result in different color loci of the phosphors in the CIE color space and that the luminous efficiency also depends greatly on the nature of the activators. Large differences can also be observed in the dominant wavelength and in FWHM. Figures 61b to 61d which

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follow show the emission spectra of these phosphors with different activators.

Figures 62a to 62e show various properties of phosphors of the invention having only europium as activator compared to phosphors having a mixture of lithium and europium as activators.

Figure 62a shows a graph with the breadth at the half maximum FWHM for a conventional phosphor of the formula $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ compared to two different phosphors of the invention $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ and $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu},\text{Li}$ for various values of x and a for Sr ($a = 2x-1$). It is clearly apparent that the conventional phosphor has a greater half-height width, while both phosphors of the invention have smaller half-height widths, it being possible to reduce the half-height width further especially by the co-doping with Li.

The thermal quenching of two phosphors of the invention is shown in figure 62b, with the integral emission intensity plotted against the temperature. Co-doping with lithium here reduces thermal quenching compared to a phosphor of the invention having only europium as the sole activator.

Figure 62c shows the comparison of the emission spectra of two phosphors of the invention, with one phosphor having been doped only with europium and the second phosphor with a mixture of europium and lithium. Both phosphors exhibit a dominant wavelength of about 604.5 nm, but the half-height width of the emission of the phosphor of the invention is reduced once again with the Eu,Li activator mixture.

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An overview of the most important spectral properties, especially the color locus in the CIE color space, the luminous efficiency LER, the dominant wavelength λ_{dom} and the half-height width of the emission FWHM is given in tabular form in figure 62d for various phosphors of the invention. As already described above, it becomes clear that the half-height width of the emission of a phosphor of the invention that has been doped with europium and lithium is lower than that of a phosphor of the invention that has been doped with europium only. In addition, the elevated quantum efficiency of the phosphor co-doped with europium and lithium is also apparent compared to the other phosphor of the invention.

The x-ray diffractograms of the phosphors $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu}$ (bottom) and $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6:\text{Eu},\text{Li}$ (top) are shown in figure 62e. An arrow highlights the characteristic x-ray diffraction reflection of the novel crystal structure of the $\text{Sr}(\text{Sr}_a\text{Ca}_{1-a})\text{Si}_2\text{Al}_2\text{N}_6$ phosphor family. Both phosphors of the invention thus have the same monoclinic crystal structure in the $P2_1$ space group.

Figures 63 to 73b show various illustrative embodiments of lighting devices 30 comprising the phosphor of the invention in cross section. The lighting devices may, for example, be white light-emitting, especially warm and/or cold white-emitting lighting devices, or red-emitting or red/orange-emitting lighting devices. These may especially be used in the automotive sector, for example as indicators or brake lights. Other possible uses are in traffic signals, in RGB applications or for "color on demand" applications, for general lighting applications, for example streetlighting or room lighting, and flash applications. These lighting devices can each be implemented with and

without a reflector dish, and multi-chip arrangements are also possible, for example for flash applications where two or more primary radiation sources are combined in one lighting device. The primary radiation sources used may, for example, be LEDs, laser diodes or else OLEDs.

Figure 63 shows a radiation source 35, for example an LED in the form of an InGaN LED disposed in a reflector dish 65 which can reflect the radiation from the radiation source 35. Disposed above this radiation source is a first matrix material 50 with a phosphor of the invention embedded therein as first phosphor 40. This phosphor is capable of absorbing the primary radiation emitted by the radiation source 35, for example blue light in the wavelength range from 300 nm to 570 nm, preferably 300 nm to 500 nm, and emitting secondary radiation in the orange/red wavelength range within a wavelength range from 570 nm to 800 nm, preferably 580 nm to 700 nm, further preferably 590 nm to 650 nm. The secondary radiation can also be emitted within a wavelength range from 580 nm to 630 nm, or 590 nm to 620 nm. Phosphors of the invention can therefore be used alone for full conversion or partial conversion of primary light to red secondary radiation, and it is also possible to use phosphors of the invention in combination with other phosphors, as described above, for production of white light, for example.

The first matrix material may comprise a multitude of materials that are transparent both to the primary radiation and to the secondary radiation that has arisen through conversion. More particularly, the first matrix material may be selected from a group of materials consisting of: glass, silicone, epoxy resin, polysilazane, polymethacrylate and polycarbonate, and combinations thereof. The polymethacrylate

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used may especially be polymethylmethacrylate (PMMA). In addition, the phosphor may also be in the form of a ceramic converter element.

Figure 64 shows a further embodiment of a lighting device in which the first phosphor 40 of the invention is disposed directly above the radiation source 35 in a separate layer. This layer may be, for example, a ceramic, a phosphor-in-glass or a silicone converter element with the first phosphor embedded therein. This embodiment too assures efficient conversion of the primary radiation to secondary radiation.

The lighting device in figure 65 is what is called a "remote phosphor" configuration in which the layer comprising the first matrix material 50 containing the first phosphor 40 of the invention is spaced apart from the radiation source 35 by an intervening space 60. In the case of such an arrangement, it is especially also possible to ensure that the conversion of the primary radiation to secondary radiation is not impaired by the heat emitted by the radiation source.

However, another alternative option is an embodiment according to figure 66 in which a volume potting material has been produced above the radiation source 35, wherein the first matrix material 35 contains a mixture of the first and second phosphor particles 40 and 45.

In addition, it is possible that the first phosphor 40 of the invention is disposed directly upon the radiation source 35, for example in the form of a silicone, phosphor-in-glass converter element, or ceramic converter element (see figure 67). The first phosphor 40 is again embedded here in a first matrix material 50 or is in the form of a ceramic converter

element. The second phosphor 45 is embedded in a potting material which surrounds the radiation source and comprises a second matrix material 55. An arrangement of this kind may be advantageous especially when the second phosphor absorbs wavelength ranges of the secondary radiation produced by the first phosphor and emits them again as light having a longer wavelength. The second matrix material here may again comprise the same materials and combinations thereof as the first matrix material. Conversely, the first phosphor may also be disposed in the surrounding potting material and the second phosphor directly upon the radiation source as converter element.

Figures 68a and 68b show alternative embodiments in which either the first or second phosphor is arranged downstream with respect to the respective other phosphor in the beam path of the primary radiation from the radiation source 35. Arrangements of this kind, and also the arrangement according to figure 67, are advantageous when the downstream phosphor is to reabsorb or convert the primary radiation produced by conversion in the upstream phosphor, or are advantageous when the upstream phosphor would absorb portions of the radiation from the downstream phosphor in the beam path of the radiation source in the case of the reverse arrangement, but this is undesirable.

Figures 69 and 70 depict various embodiments of lighting devices comprising the first phosphor 40 of the invention, in which the first matrix material 50 is disposed above the radiation source 35 either as a potting material or as a platelet. This arrangement is separated from an interference filter or filter glass 70 by means of an intervening space 60. The filter glass, which may take the form of a glass

plate, filter glass particles in a potting material or a filtering, radiation-absorbing second phosphor, may serve, for example, to absorb particular wavelength ranges of the light converted or unconverted components of the primary radiation, such that lighting devices of this kind can be used for substantially complete conversion of primary radiation to secondary radiation. In relation to combinations of a first phosphor with filter particles which may also include a second phosphor, reference is hereby made completely to German patent application DE 10 2014 105 588.8 filed on 04.17.2014.

Figures 71 and 72 show further embodiments of lighting devices comprising the phosphor of the invention, in which the layers containing the phosphors are spaced apart from one another by an intervening space 60. In each of the individual devices in these two figures, it is also possible for the first phosphors 40 and the second phosphors 45 to be interchanged.

Figures 73a and Fig. 73b each show, in cross section, a possible embodiment of a lighting device 30 suitable for flash applications. In this case, two radiation sources 35 are arranged as LED chips in a reflector dish 65 (figure 73a) or are present in two separate reflector dishes 65 (figure 73b). Phosphor particles 40 and 45 are disposed upon both LED chips, the LED chips having different phosphor mixtures/phosphors disposed in their beam paths. In this way, it is possible, for example, for two LEDs or LED modules to be present within the lighting device, which emit light of different color temperature and/or color (for example cold white and warm white or amber). A lens 75 disposed downstream of the phosphors/phosphor mixtures in the beam path serves to

mix the radiation emitted by the two LEDs or LED modules, such that an overall emission of the lighting device that results from the mixing of the radiation from the two LEDs or LED modules is perceived by an outside observer. By operating the two LEDs or LED modules with different driver currents, it is thus possible to individually adjust the overall emission emitted by the lighting device in terms of the color and/or color temperature thereof.

There follows a description of embodiments of phosphors of the invention by an alternative characterization compared to the above disclosure, but one which is consistent with the above disclosure.

A phosphor is specified. Additionally specified are a process for producing such a phosphor and a use of such a phosphor.

One problem to be addressed is that of specifying a phosphor which has comparatively narrow-band spectral emission in the red spectral region.

This problem is solved inter alia by a phosphor, by a process and by a use having the features as described herein.

In at least one embodiment, the phosphor is set up to emit red light. Red light means that the phosphor emits radiation having a dominant wavelength between 585 nm and 640 nm inclusive, especially between 590 nm and 615 nm inclusive.

The dominant wavelength is especially that wavelength which is found to be the point of intersection of the spectral

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color line of the CIE standard chromaticity diagram with a straight line, this straight line proceeding from the white point in the CIE standard chromaticity diagram and running through the actual color locus of the radiation. In general, the dominant wavelength differs from a wavelength of maximum intensity. More particularly, the dominant wavelength in the red spectral region is at smaller wavelengths than the wavelength of maximum intensity.

In at least one embodiment, the phosphor has the general empirical formula $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$. It is possible that the phosphor includes further elements, for instance in the form of impurities, in which case these impurities taken together preferably have a maximum proportion by weight in the phosphor of not more than 0.1 permille or 10 ppm, parts per million.

In at least one embodiment of the phosphor, $x > 0.8$ or $x \geq 0.82$ or $x \geq 0.85$ or $x \geq 0.89$. Alternatively or additionally, $x \leq 1$ or $x < 1$ or $x \leq 0.98$ or $x \leq 0.95$ or $x \leq 0.92$.

In at least one embodiment, the proportion of the Sr lattice sites which have been replaced by Eu is at least 0.01% or 0.1% or 0.35% or 0.5%. Alternatively or additionally, this proportion is at most 10% or 5% or 3% or 2.2% or 1.8%.

In at least one embodiment, the phosphor, in an x-ray structure analysis, has a reflection having the Miller indices $1\bar{2}1$ when an orthorhombic description is used as the basis of the crystallographic cell. This statement includes descriptions of equivalent symmetry such as $1\bar{2}\bar{1}$.

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In at least one embodiment, the phosphor is set up to emit red light and preferably to be excited by blue light and has the general empirical formula $Sr_xCa_{1-x}AlSiN_3:Eu$ with $0.8 < x \leq 1$. A proportion of the Sr lattice sites between 0.1% and 5% inclusive has been replaced by europium. In an x-ray structure analysis, the phosphor in orthorhombic description has a reflection having the Miller indices $1\bar{2}1$.

Phosphors which can be excited in the ultraviolet spectral range up to the blue/green spectral range and which emit red light are of great significance for the production of white light-emitting diodes. Specifically in the case of light-emitting diodes having low color temperatures, called warm white light-emitting diodes, and/or in the case of light-emitting diodes having a high color rendering index, phosphors of this kind are required. Phosphors of this kind are also required in a multitude of other applications, for instance for display backlighting, what are called color-on-demand applications or else for orange and red full conversion light-emitting diodes. Use in combination with an organic light-emitting diode, OLED for short, is likewise possible. The phosphor described here is usable for such applications.

In at least one embodiment, the phosphor in a powder diffractogram on irradiation with monochromatic $Cu-K_{\alpha 1}$ radiation has a reflection at an angle 2θ between 36.7° and 37.0° , according to the composition of the phosphor. The exact position of this reflection depends on the parameter x in the general empirical formula of the phosphor. An intensity of this reflection, especially based on a main reflection, is preferably at least 0.3% or 0.5% and/or at most 10% or 8% or 5% or 4%.

In at least one embodiment, the dominant wavelength of the phosphor is at least 596 nm or 598 nm. Alternatively or additionally, the dominant wavelength is at most 606 nm or 604 nm. The wavelength of maximum intensity is, for example, at least 605 nm or 610 nm and/or at most 630 nm or 625 nm.

In at least one embodiment, the phosphor has a spectral half-height width at half the maximum height, FWHM or full-width at half maximum for short, of at least 70 nm or 75 nm or 78 nm. The maximum of this spectral range is preferably 90 nm or 87 nm or 84 nm or 82 nm.

In at least one embodiment, the phosphor has an absorption maximum in the blue spectral region, especially a relative absorption maximum. The blue spectral region especially refers to wavelengths of at least 400 nm and/or of at most 480 nm. For example, the absorption maximum is at at least 410 nm or 420 nm and/or at at most 450 nm or 440 nm.

The abovementioned values relating to the spectral properties of the phosphor especially apply at room temperature, i.e. at about 300 K.

Additionally specified is a process for producing such a phosphor. Features of the phosphor are therefore also disclosed for the process, and vice versa.

In at least one embodiment, the process has at least the following steps, preferably in the sequence specified:

- A) providing reactants in the solid state for Sr, Al, Si and Eu and optionally for Ca,
- B) mixing the reactants,

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- C) heating the reactants under a forming gas atmosphere to at least 1500°C and forming a calcined cake, and
- D) comminuting the calcined cake to give the phosphor.

In at least one embodiment of the process, at least step C) or else all steps are effected at atmospheric pressure. More particularly, the process in that case is not effected under high pressure conditions. Preferably, the atmospheric pressure and/or a total pressure is between 0.9 bar and 1.5 bar or 0.95 bar and 1.05 bar inclusive.

Reactants and sources used for strontium, aluminum and/or calcium may be the respective pure metals or else metal alloys with the appropriate metals. Reactants used may likewise be silicides, nitrides, oxynitrides, halides and/or oxides of these metals. In addition, it is possible to use mixtures of these compounds.

Reactants or sources used for silicon for the production of the phosphor may be a silicon-metal compound, a silicon nitride, an alkaline earth metal silicide, silicon diimide, or a mixture of these compounds. Preference is given to using silicon nitrides and/or silicon metals.

Reactants or sources used for Eu may be metallic europium, a europium alloy, a europium oxide, a europium nitride or a europium halide. It is likewise possible to use mixtures of these compounds. Preference is given to using europium oxide as reactant for europium.

In at least one embodiment, a melting agent and/or a flux is used for the improvement of crystallinity and/or to promote crystal growth. For this purpose, preferably chlorides,

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fluorides, halides and/or boron compounds of the alkaline earth metals used are employed. Combinations of two or more melting agents or fluxes may also be used. Melting agents and/or fluxes used are especially at least one of the following substances: LiF, LiCl, NaF, NaCl, SrCl₂, SrF₂, CaCl₂, CaF₂, BaCl₂, BaF₂, NH₄Cl, NH₄F, KF, KCl, MgF₂, MgCl₂, AlF₃, H₃BO₃, B₂O₃, Li₂B₄O₇, NaBO₂, Na₂B₄O₇, LiBF₄.

In at least one embodiment, the starting substances, especially for Sr, Ca, Al and/or Si and also Eu, are weighed out according to the general empirical formula of the phosphor. It is possible that the alkaline earth metal components are also weighed out with an excess, in order to compensate for any evaporation losses that occur during the synthesis.

In at least one embodiment, step D) is followed by a step E). In step E), the phosphor is calcined further, which can also be referred to as heat treatment. The calcination is especially effected at a temperature of at least 1500°C and preferably under a nitrogen atmosphere or forming gas atmosphere. Forming gas refers to a mixture of N₂ and H₂. The temperature of at least 1500°C in steps C) and/or E) is preferably employed for at least four hours or six hours. For example, in each of steps C) and E), a temperature of 1650°C ± 50°C is employed.

In at least one embodiment, the reactants are mixed in a ball mill or in a tumbling mixer. In the mixing operation, it may be advantageous to choose the conditions such that a large amount of energy is introduced into the mixture, which results in grinding of the reactants. The resultant increase in homogeneity and reactivity of the mixture can have a

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positive influence on the properties of the resulting phosphor.

By controlled alteration of the bulk density or by modification of the agglomeration of the reactant mixture, it is possible to reduce the formation of secondary phases. In addition, a particle size distribution, a particle morphology and/or a yield of the resulting phosphor can be affected. Techniques of particular suitability for the purpose are sieving and pelletizing operations, including use of suitable additives.

In at least one embodiment, a tempering is effected, especially in a crucible made from tungsten, molybdenum or boron nitride. The tempering is preferably effected in a gas-tight oven in a nitrogen atmosphere or in a nitrogen/hydrogen atmosphere. The atmosphere may be flowing or stationary. It is additionally possible for carbon in finely divided form to be present in the oven space. Also possible are multiple tempering steps of the phosphor, in order to improve the crystallinity or particle size distribution. Further advantages may be a lower defect density, associated with improved optical properties of the phosphor and/or a higher stability of the phosphor. Between the tempering steps, the phosphor may be treated in a wide variety of different ways, or it is possible to add substances such as melting agents to the phosphor.

For grinding of the phosphor, it is possible, for instance, to use a mortar mill, a fluidized bed mill or a ball mill. In the grinding operation, it is to be ensured that the proportion of splintered grains produced is kept to a

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minimum, since these can worsen the optical properties of the phosphor.

The phosphor can additionally be washed. For this purpose, the phosphor can be washed in water or in aqueous acids such as hydrochloric acid, nitric acid, hydrofluoric acid, sulfuric acid, organic acids or a mixture of these. The phosphor may alternatively or additionally be washed in an alkali such as sodium hydroxide solution, potassium hydroxide solution, an aqueous ammonia solution or mixtures thereof. Alternatively or additionally, washing in organic solvents such as acetone, propanol and/or phenol is possible. The washing preferably follows after the grinding.

In at least one embodiment, for instance, the tempering, further calcining, grinding, sieving and/or washing result in removal of secondary phases, glass phases or other contamination and hence an improvement in the optical properties of the phosphor. It is also possible by this treatment to selectively remove or dissolve small phosphor particles and to optimize the particle size distribution for the application. In addition, such a treatment can alter a surface of the phosphor particles in a controlled manner, for example the removal of particular constituents from the particle surface. This treatment can, also in conjunction with a downstream treatment, lead to improved stability of the phosphor.

Additionally specified is the use of such a phosphor. Features relating to use are therefore also disclosed for the process and the phosphor, and vice versa.

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In at least one embodiment, the phosphor is used in a light-emitting diode. The light-emitting diode comprises at least one semiconductor chip that emits in the blue spectral region in operation. The phosphor is arranged downstream of the semiconductor chip along a beam path.

The blue light produced by the semiconductor chip is partly or fully absorbed by the phosphor and converted to red light. It is possible that further phosphors are present, especially for generation of green and/or yellow light. In addition, the light-emitting diode preferably emits mixed radiation including blue light from the semiconductor chip and converted radiation from the phosphor, and also green and/or yellow light from the further phosphor.

A phosphor described here is elucidated in detail hereinafter by embodiments with reference to the drawing. Identical reference numerals indicate identical elements in the individual figures. However, the drawings are not to scale; instead, individual elements may be shown in excessively large size for better understanding.

The figures show:

figure 74 starting weights for embodiments and for modifications of phosphors described here and the color locus that they emit,

figures 75 to 87 schematic diagrams of properties of phosphors described here on excitation with blue light,

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figures 88 to 90 data from an x-ray structure analysis of phosphors described here,

figure 91 a schematic diagram of a structure of a phosphor described here, and

figure 92 a schematic diagram of a structure of a modification of a phosphor described here.

One embodiment of a phosphor described here can be prepared as follows:

Reactants used for the synthesis of the phosphor of the general empirical formula $Sr_xCa_{1-x}AlSiN_3:Eu$ are the binary nitrides of the constituent elements, i.e. Ca_3N_2 , Sr_3N_2 , AlN and Si_3N_4 . Since these are highly oxidation-sensitive and hydrolysis-sensitive substances, what is called a glovebox is employed, under an N_2 atmosphere with $O_2 < 1$ ppm and $H_2O < 1$ ppm. In addition, doping with Eu^{2+} is accomplished using Eu_2O_3 . The reactants are weighted out such that the following atomic ratio is present:

$Ca : Sr : Al : Si : Eu = (1-x) : x : 1 : 1 : y$, where y corresponds to the degree of doping, i.e. the proportion of Sr lattice sites which are substituted by Eu. In addition, various fluxes are added; see the table in figure 74. A reactant mixture is scaled up to a total starting weight of 50-100 g with retention of the atomic ratios described above; see the table in figure 74 likewise.

The reactant mixture is introduced into a PET mixing vessel together with ZrO_2 balls and mixed on a roller table in a glovebox for 6 h. Subsequently, the balls are removed from

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the mixture and the powder is transferred into a closed molybdenum crucible. This crucible is placed into an outer tungsten crucible, a semicircular open tungsten tube, and transferred into a tube furnace. There is a flow of 3 L/min of forming gas with 92.5% N₂ and 7.5% H₂ through the tubular furnace during the run time. In the tubular furnace, the mixture is heated at a rate of 250 K/h to 1650°C, kept at this temperature for 4 h and then cooled down at 250 K/h to 50°C. The calcined cake formed is taken out after the furnace has cooled, comminuted with a mortar mill and sieved through a sieve having a mesh size of 31 µm. The sieve fraction < 31 µm is the phosphor used.

The sieving may optionally be followed by a further calcining, tempering and/or washing operation.

Illustrative starting weights m in g and resulting color loci CIE x, CIE y, also referred to as chromaticity coordinates, of the emission spectrum of the particular phosphor in the CIE standard chromaticity diagram on excitation with blue light at 460 nm and on complete absorption of the blue light are also listed in the table in figure 74. For each of the embodiments in the table, 0.8 ≤ x ≤ 1.

Figures 75 to 78 show the properties of radiation emitted by the phosphor.

Figure 75 shows emission spectra and figure 76 reflectance spectra of Sr_xCa_{1-x}AlSiN₃:Eu phosphors. The wavelength λ is plotted against the intensity I and the reflectivity R. The emission spectra show an unexpectedly narrow spectral emission of the phosphor with x = 0.9. At the same time, the

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phosphor with $x = 0.9$ features a strong absorption; see figure 3. The absorption is found to be approximately $1-R$.

Figure 77 shows a dependence of a spectral half-height width FWHM of the emission on the Sr content, i.e. the parameter x in $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$. Up to an Sr content of 80%, i.e. $x = 0.8$, a very small change in the half-height width FWHM with rising x is observed. Surprisingly, from an Sr content of > 80% onward, an abrupt decrease in the half-height width FWHM is suddenly observed.

Figure 78, analogously to figure 77, shows the spectral half-height width FWHM as a function of the parameter x . Also stated are a dominant wavelength λ_{dom} of the spectrum emitted by the phosphor and the Eu content. Surprisingly, a phosphor having 90% Sr, with a comparable dominant wavelength λ_{dom} , has a much smaller half-height width FWHM compared to conventional phosphors having only a relatively low Sr content. The abrupt decrease in the half-height width FWHM is virtually independent of the Eu content used in the sample.

The phosphor with $x \geq 0.8$ thus features a small half-height width FWHM of the emission and a very high luminous efficiency LER with simultaneously high internal quantum efficiency QI and external quantum efficiency QE; see the table in figure 79. In addition, a relative brightness B is stated. To calculate the external quantum efficiency QE, the mean reflectance within the range from 450 nm to 470 nm was employed; measurement was effected in pressed powder tablets at an excitation wavelength of 460 nm.

Figure 80 shows a comparison of conversion efficiencies of various warm white light-emitting diodes, LEDs for short. A

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mixture of two phosphors was used in each case, with the green light-emitting phosphor G remaining the same and the red-emitting phosphor R being varied. Stated on the abscissa axis is the type of red-emitting phosphor R. The ordinate axis states the relative efficiency E. The phosphors were excited with a blue-emitting semiconductor chip having a dominant wavelength of 446 nm.

All phosphor mixtures were adjusted so as to achieve, in the CIE standard chromaticity diagram, a color locus close to Planck with a correlated color temperature CCT of about 2700 K. The color rendering index CRI of all LEDs measured is 80 ± 1 . All red phosphors R used show a comparable dominant wavelength of about 600.5 nm \pm 1 nm.

Further details of the phosphor mixtures as shown in figure 80 can also be found in the table in figure 81. Additionally stated are the relative efficiency E, a phosphor concentration c and a ratio V of the green phosphor G and the red phosphor R.

Figure 82 shows a comparison of conversion efficiency and color rendering index for various warm white light-producing LEDs. A mixture of two phosphors was used in each case, with the green phosphor G being kept constant and the red phosphor R being varied, analogously to the table in figure 81. All phosphor mixtures were adjusted so as to achieve a color locus close to Planck with a correlated color temperature CCT of about 2700 K. The efficiency E, illustrated by the bar, of a warm white light-producing LED comprising the novel phosphor with 90% Sr shows a much higher efficiency and simultaneously an improved color rendering index CRI,

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symbolized as rhombuses, compared to LEDs comprising a red phosphor with only 80% Sr.

Further data relating to the LED measurements from figure 82 can be found in the table in figure 83, analogously to the table in figure 81. The efficiency E of a warm white light-producing LED at correlated color temperature CCT of about 2700 K comprising a novel red phosphor with 90% Sr is distinctly higher here as well and, moreover, an elevated color rendering index CRI is achievable.

Red phosphors composed of the $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ material system were subjected to a hydrolysis test, in order to assess an aging stability of the phosphor with respect to air humidity; see figure 84. For this purpose, the corresponding phosphor powders were stored at 130°C and 100% relative air humidity for 40 h. The absorption A of the phosphors in the blue spectral region between 450 nm and 470 nm was measured both before and after this treatment. A measure of the stability of a phosphor with respect to hydrolysis, i.e. the decomposition of the phosphor in the presence of water, is considered to be the decrease in the absorption capacity in the blue spectral region. With increasing Sr content, a significant increase in hydrolysis sensitivity is at first observed. Surprisingly, however, the novel phosphor with 90% Sr is more hydrolysis-stable than a phosphor with only 80% Sr content.

Figure 85 shows thermal quenching characteristics of two red phosphors compared to one another. The two phosphors have a comparable emission color with a dominant wavelength of around 600 nm. Surprisingly, the novel phosphor having high Sr content, in spite of a higher Eu content, has a smaller

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decrease in emission intensity I with rising temperature compared to the reference phosphor.

Figure 86 illustrates the relative emission intensity I as a function of the content of Eu as activator. The Eu content is reported in percent.

With regard to phosphors composed of the $\text{CaAlSiN}_3:\text{Eu}$ system, the literature reports that, with rising activator content, especially $> 0.8\%$ Eu, a conversion efficiency stagnates; see table 1 in EP 2 135 920 A1.

Surprisingly, the novel phosphor with a high Sr content shows different behavior. With rising Eu content, the emission intensity I, even in the case of an Eu content $> 1\%$, continues to increase in an approximately linear manner. This property offers various technical advantages for application, especially lower phosphor demand and the possibility of achieving color loci with greater CIE x.

Figure 87 shows a dependence of the dominant wavelength λ_{dom} of the emission on the activator content y for the novel phosphor $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$ with $X = 0.9$. With rising activator content, the luminescence signal moves toward higher wavelengths in an approximately linear manner. This allows, for example, the color rendering index CRI of a warm white light-emitting LED to be increased; see also the LED embodiments according to figures 80 to 82.

Figure 88 shows an x-ray powder diffractogram of the phosphor $\text{Sr}_{0.8}\text{Ca}_{0.2}\text{AlSiN}_3:\text{Eu}$, which has been produced by means of the synthesis described here. Surprisingly, the phosphor produced from Sr nitride, Ca nitride, AlN, Si_3N_4 and Eu_2O_3 under

- 110 -

atmospheric pressure consists of a pure phase by x-ray crystallography. No reflections of secondary phases such as AlN or $(\text{Sr,Ca})_2\text{Si}_5\text{N}_8$ are observed.

Figure 89 shows x-ray powder diffractograms of phosphors $\text{Ca}_{1-x}\text{Sr}_x\text{AlSiN}_3:\text{Eu}$ with various values of x , which have been prepared by means of the synthesis described here. From a substitution level of $x > 0.8$ upward, the occurrence of an additional reflection R at 2θ of 36.7° to 37.0° is observed. This reflection cannot be explained by the structure model of known $(\text{Sr,Ca})\text{AlSiN}_3$. Nor was it possible to assign the reflection R to any compounds from databases.

In orthorhombic description, this reflection R originates from the $1\bar{2}1$ lattice plane. The exact position of this reflection depends on the substitution level x . If it were a reflection R from any unidentified extraneous phase, this shift would not be expected.

In order to describe the structure of the novel phosphor described here, the following steps were conducted in order. A summary of the results, important R values and the fundamental refined parameters can be found in the table in figure 90.

1) A Rietveld refinement was conducted with the known phases $\text{Sr}_{0.99}\text{Eu}_{0.01}\text{AlSiN}_3$ from ICSD 98-041-9410, AlN from ICSD 98-060-8626 and SrF_2 from ICSD 98-004-1402. The crystal structure data of $\text{Sr}_{0.99}\text{Eu}_{0.01}\text{AlSiN}_3$ were fitted as $\text{Ca}_{0.1}\text{Sr}_{0.89}\text{Eu}_{0.01}\text{AlSiN}_3$.

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2) All reflections were fitted by means of a profile parameter fit, with equal FWHM for all reflections. Any reflections which were assigned to extraneous phases, for instance SrF_2 and AlN , were eliminated from the search. The other reflections were used for a lattice parameter search. This lattice parameter search found that almost all reflections can be described with the original cell, but without quenching conditions. For this reason, in the next step, a Rietveld refinement was conducted with the original structure, but transferred to the $\text{P}1$ space group.

3) A trial refinement of the experimental data was conducted on the basis of the structure model for $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3$ known from the literature, except that it had been transferred to the lower-symmetry $\text{P}1$ space group; see also the preceding step 2. This refinement likewise converged, but does not explain the reflection R observed.

4) In order to explain the additional reflections observed, a new structure model different than the known CaAlSiN_3 structure was set up. The new structure model of the phosphor described here is distinctly different than the CaAlSiN_3 structure already known. In the crystallographic sense, this is a superstructure variant. This structure can be derived formally from that of CaAlSiN_3 by a reduction in symmetry. In the structure model thus derived for the novel phosphor, there is a good explanation and description of the at least one reflection R additionally observed.

This structure model of the novel phosphor differs from the structure model from the above-elucidated step 3). In step 3), the known structure model of CaAlSiN_3 was merely

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described in an alternative, lower-symmetry space group. Only through the actual introduction of a new structure model different than CaAlSiN_3 is a good description of the experimentally observed reflections possible, especially of the new reflection R.

For this purpose, it is necessary in particular to split up the position having a mixed Sr/Ca population, which has a mixed population in the original $\text{Cmc}2_1$ space group and describes four alkaline earth metal atoms simultaneously for reasons of symmetry, into four individual positions. In the model of the conventional phosphor, all four positions have mixed populations of Sr and Ca. In the model of the novel phosphor, three of the positions are populated only by Sr, and just one of the positions has mixed population both with Sr and with Ca.

The new reflection R shown is thus a superstructure reflection which can be described in $\text{P}1$, but not in $\text{Cmc}2_1$, since it infringes the quenching conditions for this space group.

The refinement of the powder x-ray data observed, based on the known structure model in space group $\text{Cmc}2_1$, leads to the goodness factors in the first column in figure 17, corresponding to the above step 1. An alternative description of the same known structure model in the lower-symmetry $\text{P}1$ space group leads to comparable goodness factors, reported in the third column in figure 90, corresponding to the above step 3. Only with the aid of the description with the novel structure model different than CaAlSiN_3 are a complete description of all the reflections observed and hence

significantly improved goodness factors achieved, corresponding to the above step 4.

Figure 91 shows a schematic perspective view of the structure model of the novel phosphor with $x \geq 0.8$. The positions shown in a dark color are populated only by Sr. The position shown in white color has a mixed Ca/Sr population.

In comparison, the perspective diagram according to figure 92 illustrates the structure of the CaAlSiN_3 phosphor with small x in the $\text{Cmc}2_1$ space group. Shown in a dark color are the four positions having a mixed Ca/Sr population.

The novel phosphor described here offers the following particular advantages:

- lower half-height width of emission, associated with higher luminous efficiency at the same dominant wavelength,
- the possibility of achieving higher activator concentrations of Eu at $> 0.8\%$ with simultaneously high quantum efficiency and conversion efficiency, associated with a lower phosphor requirement in LED applications and simplified processibility,
- improved aging stability with respect to moisture compared to conventional $(\text{Sr,Ca})\text{AlSiN}_3:\text{Eu}$ having a low Sr content, and
- improvement in the thermal stability.

The invention described here is not restricted by the description with reference to the embodiments. Instead, the

invention encompasses every new feature and every combination of features.

CLAIMS:

1. A phosphor for emission of red light, having the general empirical formula $\text{Sr}_x\text{Ca}_{1-x}\text{AlSiN}_3:\text{Eu}$

5 wherein $0.8 < x \leq 1$,

wherein between 0.1% and 5% inclusive of the Sr, Ca and/or Sr/Ca lattice sites have been replaced by Eu,

wherein a full-width at half maximum is between at least 70 nm and at most 90 nm, and

10 wherein, in an x-ray structure analysis, the phosphor in orthorhombic description exhibits a reflection (R) having the Miller indices **121**.

2. The phosphor according to claim 1,

15 wherein $0.85 \leq x \leq 0.95$, and

wherein between 0.35% and 2.2% inclusive of the Sr lattice sites have been replaced by Eu.

3. The phosphor according to claims 1 or 2,

20 which exhibits, in a powder diffractogram on monochromatic irradiation with $\text{Cu}-K_{\alpha 1}$ radiation, the reflection (R)

having the Miller indices **121** at 2θ between 36.7° and 37.0° .

25 4. The phosphor according to claim 3,

in which the reflection (R), based on a main reflection, exhibits an intensity between 0.3% and 8% inclusive.

5. The phosphor according to any one of claims 1 to 4,

which has a dominant wavelength between 596 nm and 606 nm inclusive,

wherein a breadth of a radiation spectrum emitted by the phosphor at half the maximum height is between 75 nm and 5 87 nm inclusive, and

wherein the phosphor has a relative absorption maximum in the wavelength range between 410 nm and 450 nm inclusive and can be induced to emit radiation with blue light.

10 6. The phosphor according to any one of claims 1 to 5, wherein the phosphor has a structure comprising $(\text{Si}/\text{Al})\text{N}_4$ tetrahedra arranged in a 3D network, in which layers in an a-c plane are linked in a b-direction, wherein intercalated between the network, layer by layer 15 in each case, are pure Sr positions and positions having a mixed Sr/Ca population.

7. A process for producing a phosphor according to any one of claims 1 to 6, 20 having the steps of:
A) providing reactants in the solid state for Sr, Al, Si and Eu and optionally for Ca,
B) mixing the reactants,
C) heating the reactants under a nitrogen atmosphere or 25 under a forming gas atmosphere to at least 1500°C and forming a calcined cake, and
D) comminuting the calcined cake to give the phosphor.

8. The process according to claim 7,

in which the reactants for Sr, Al and/or Ca is each a pure metal, a metal alloy, a silicide, a nitride, an oxynitride, an oxide, a halide, or a mixture thereof, wherein the reactant for Si is a silicon metal, a silicon nitride, an alkaline earth metal silicide, a silicon diimide, or a mixture thereof, and wherein the reactant for Eu is at least one of an europium metal, an europium oxide, an europium nitride, and an europium halide.

10

9. The process according to claim 7 or 8, wherein heating the reactants comprises adding one or more of the following substances as a melting agent and/or as a flux: LiF, LiCl, NaF, NaCl, SrCl₂, SrF₂, CaCl₂, CaF₂,

15

BaCl₂, BaF₂, NH₄Cl, NH₄F, KF, KCl, MgF₂, MgCl₂, AlF₃, H₃BO₃, B₂O₃, Li₂B₄O₇, NaBO₂, Na₂B₄O₇, and LiBF₄.

20

10. The process according to any one of claims 7 to 9, further comprising, after comminuting the calcined cake, calcinating the phosphor at a temperature of at least 1500°C under forming gas atmosphere.

25

11. A light-emitting diode comprising:
at least one semiconductor chip configured to emit blue light;
a phosphor arranged downstream of the semiconductor chip along a beam path wherein the phosphor emits red light during operation and has the general empirical formula:
Sr_xCa_{1-x}AlSiN₃:Eu
wherein x is 0.8 < x ≤ 1,

wherein between 0.1% and 5% inclusive of the Sr, Ca and/or Sr/Ca lattice sites have been replaced by Eu,
wherein a full-width at half maximum is between at least
70 nm and at most 90 nm, and
5 wherein the phosphor in orthorhombic description exhibits
a reflection (R) having the Miller indices **1 $\bar{2}$ 1**.

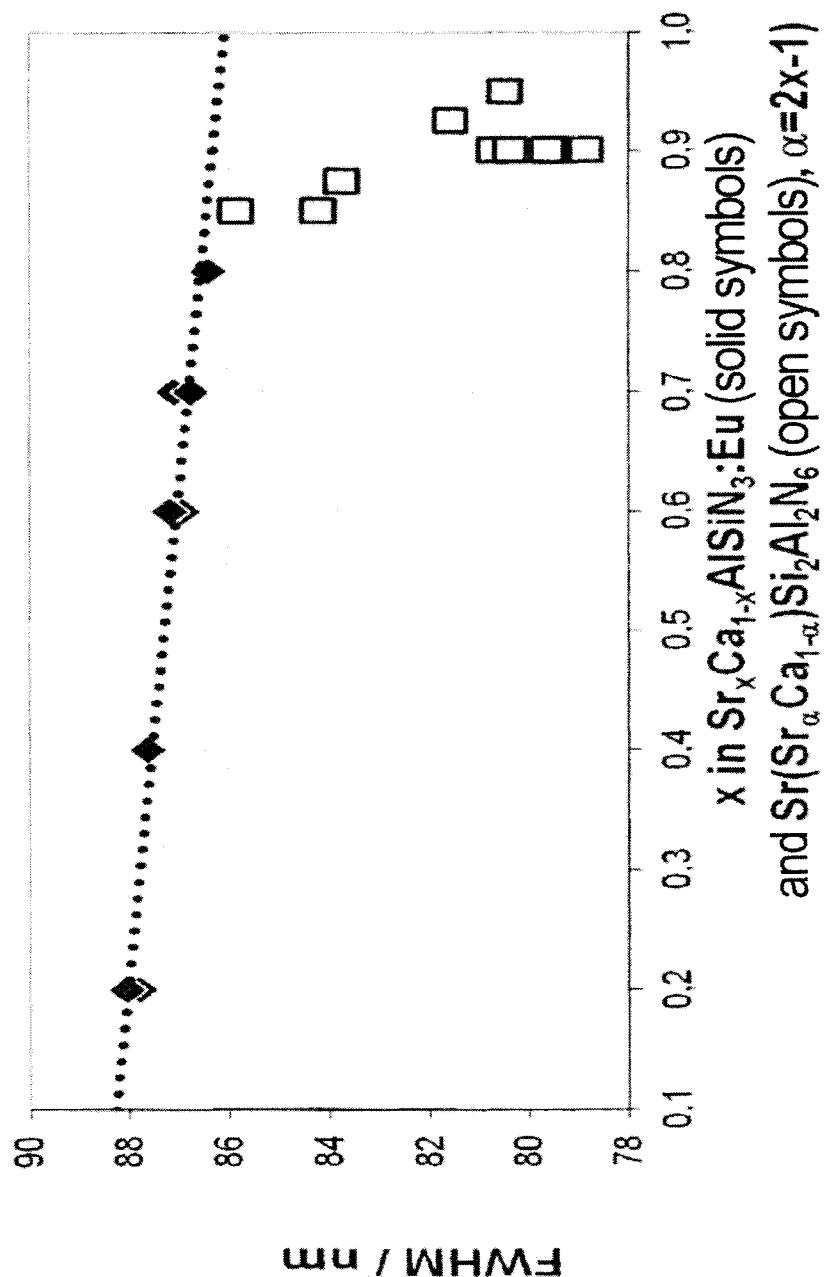


Fig. 1

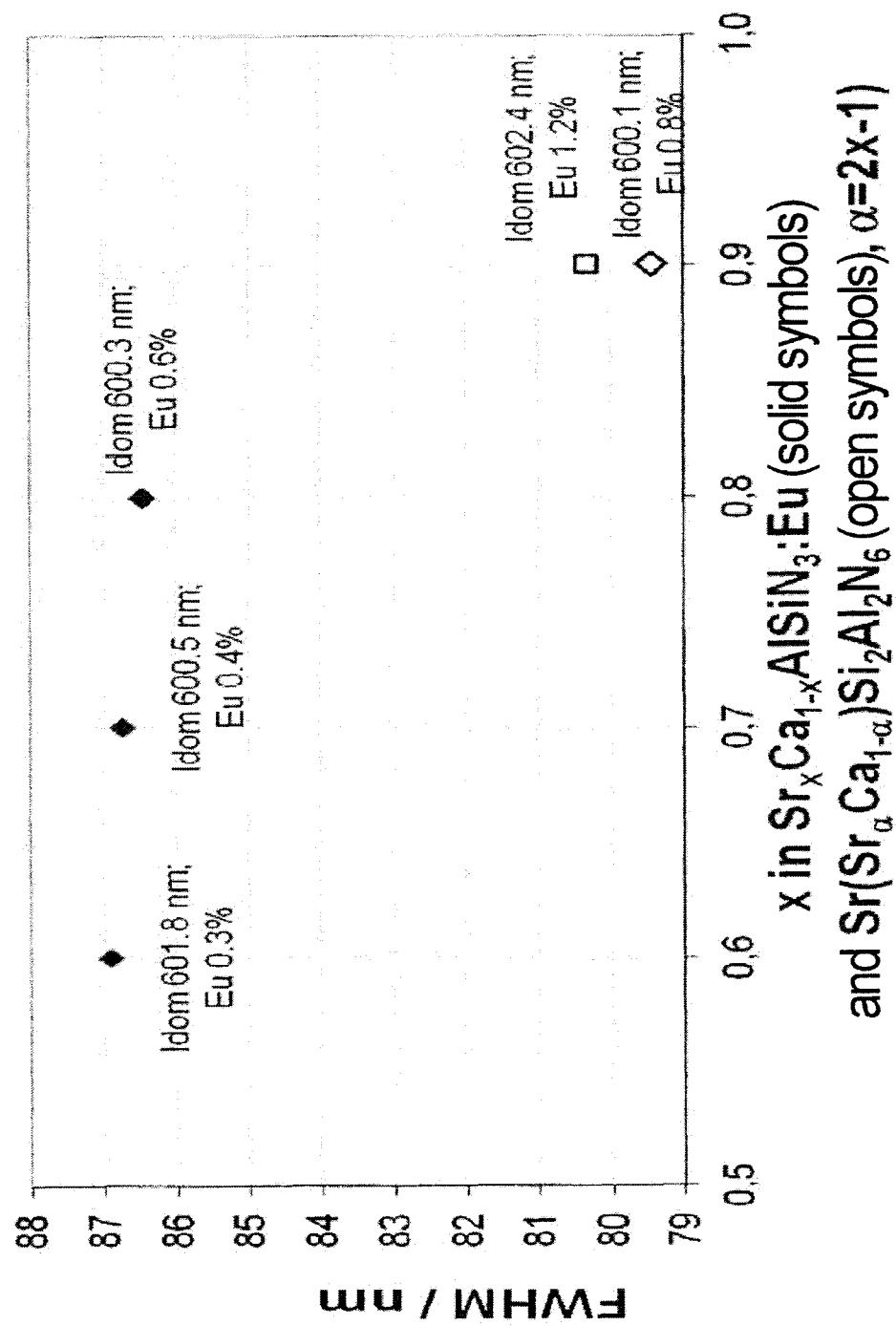


Fig. 2

Fig. 3

	x	y	Idom/nm	FWHM/nm	QI	QE	LER	B
(Sr _{0.7} Ca _{0.3})AlSiN ₃ :Eu (0.4%)	0.628	0.370	600,5	86,7	100%	100%	100%	100%
(Sr _{0.8} Ca _{0.2})AlSiN ₃ :Eu (0.6%)	0.628	0.371	600,3	86,5	99%	110%	101%	111%
Sr(Sr _{0.8} Ca _{0.2})Si ₂ Al ₂ N ₆ :Eu (0.8%)	0.626	0.372	600,1	79,5	99%	113%	110%	125%
(Sr _{0.6} Ca _{0.4})AlSiN ₃ :Eu (0.4%)	0.639	0.359	603,1	87,2	100%	100%	100%	100%
Sr(Sr _{0.8} Ca _{0.2})Si ₂ Al ₂ N ₆ :Eu (1.2%)	0.637	0.362	602,4	80,4	99%	113%	111%	126%

Fig. 4

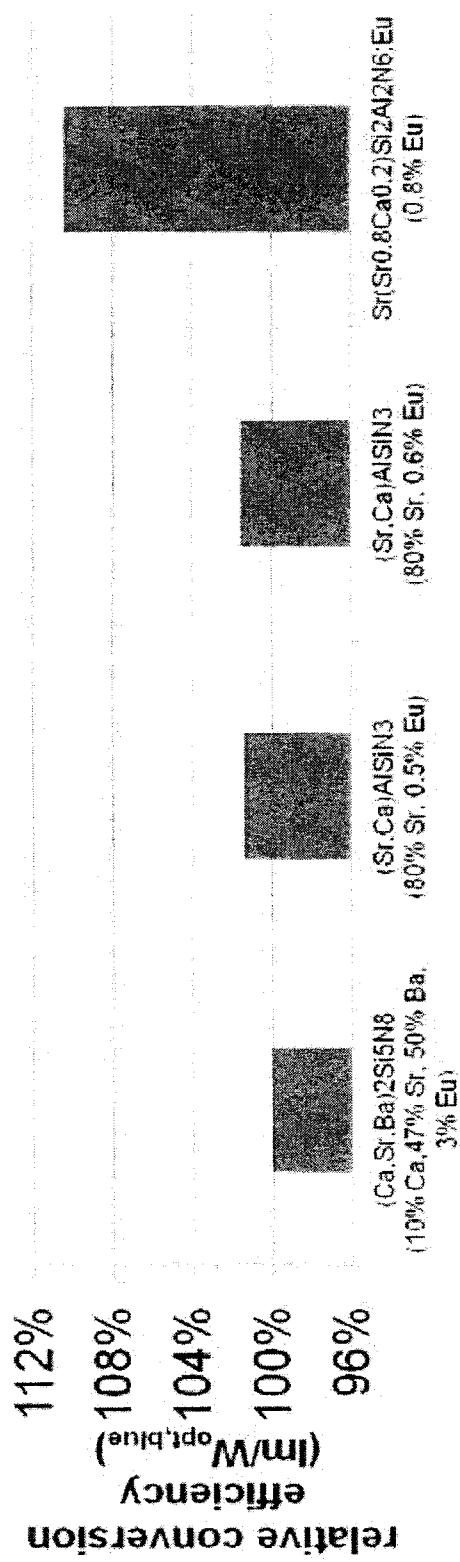


Fig. 5

comment	Blue LED $\lambda_{\text{dom}} / \text{nm}$	Green Phosphor $(\text{Lu,Ce})(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	Red Phosphor $(\text{Ca},\text{Sr},\text{Ba}_2\text{Sr}_6\text{N}_3$ (10% Ca, 41% Sr, 50% Ba, 3% Eu)	$\lambda_{\text{em}} / \text{nm}$ (red phosphor concentrati on)	phosphor ratio green/ red	CIE x	CIE y	CCT / K	CRI	rel. eff.
Comparative example 1	446	$(\text{Lu,Ce})(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$(\text{Sr,Ca})\text{Al}(\text{SiN}_3$ (80% Sr, 15% Eu))	601,5	15,5%	4,45:1	0,461	0,408	2669	79 100,0%
Comparative example 3	446	$(\text{Lu,Ce})(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$(\text{Sr,Ca})\text{Al}(\text{SiN}_3$ (80% Sr, 0,6% Eu))	599,4	16,0%	2,5:1	0,459	0,416	2754	80 101,4%
Comparative example 4	446	$(\text{Lu,Ce})(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$(\text{Sr,Ca})\text{Al}(\text{SiN}_3$ (80% Sr, 0,6% Eu))	600,3	16,8%	2,6:1	0,461	0,408	2655	81 101,5%
Inventive example 1	446	$(\text{Lu,Ce})(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$\text{Sr}(\text{Sr}_{10},\text{Ce}_{0,2})\text{Al}_2\text{Si}_2\text{N}_6\text{-Eu}$ (0,8% Eu)	600,1	16,7%	3,7:1	0,462	0,411	2681	79 110,4%

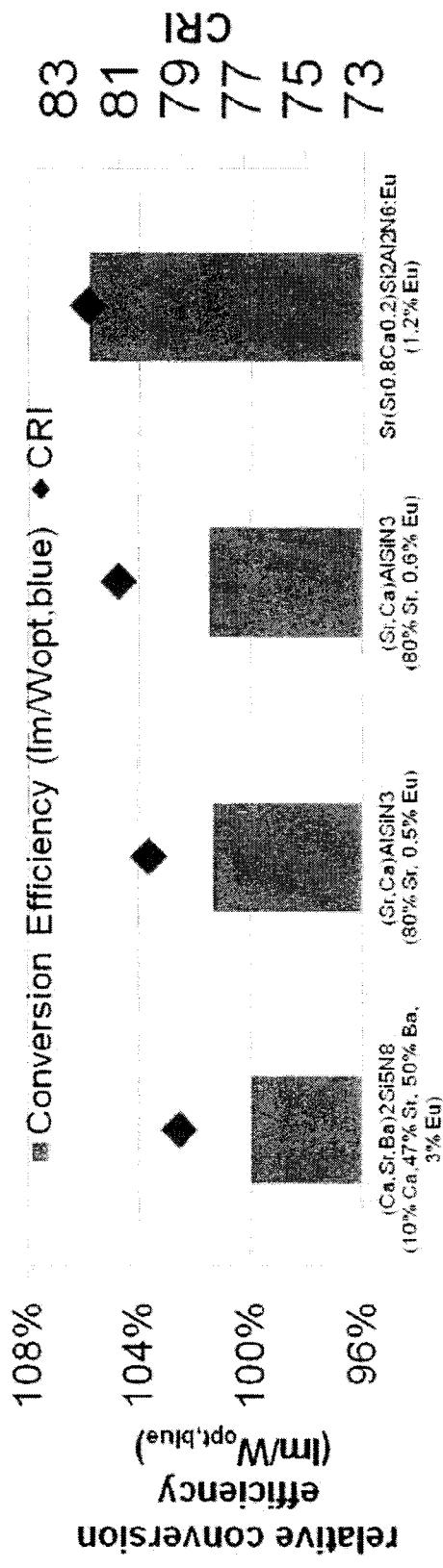
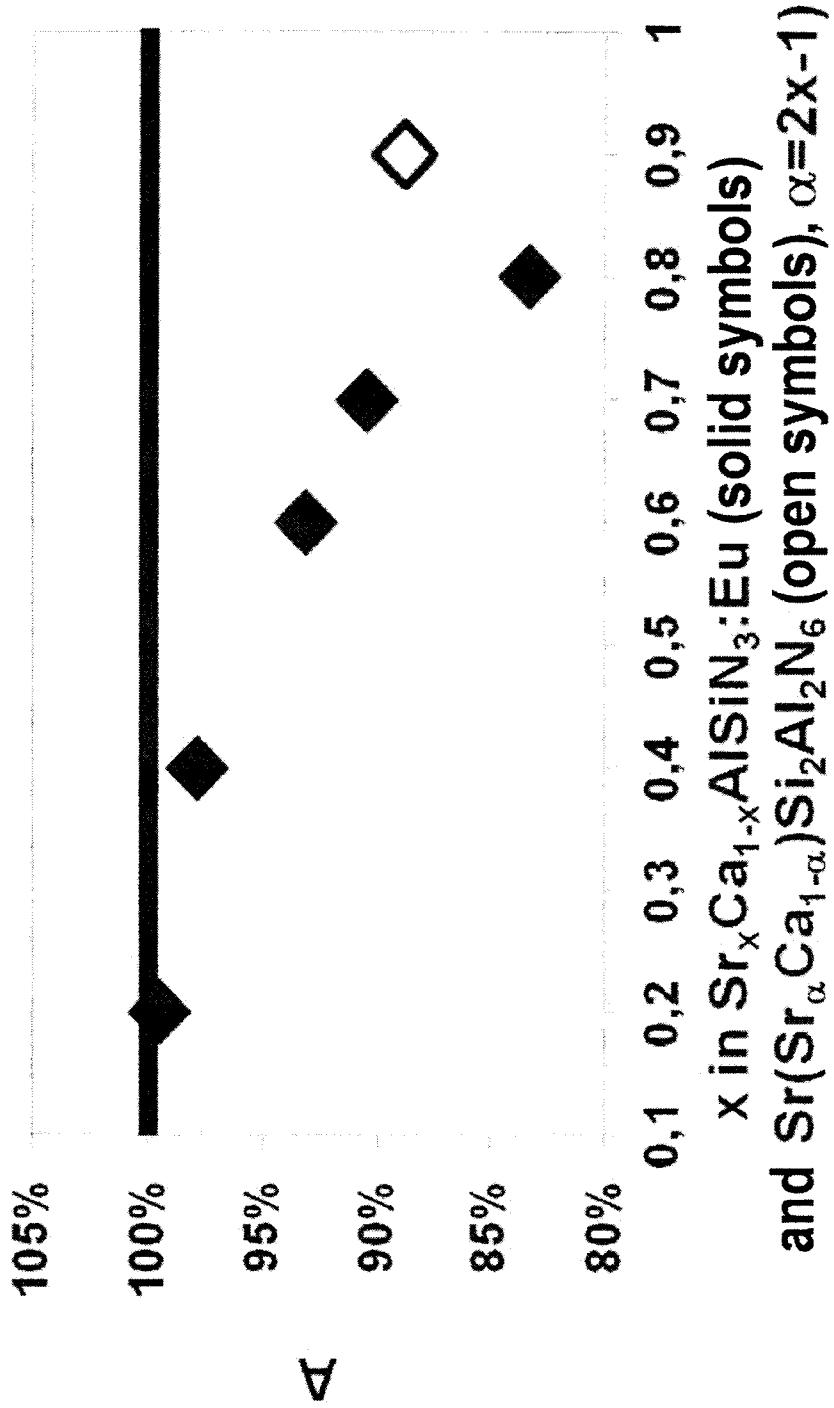


Fig. 6

Fig. 7

comment	Blue LED $\lambda_{460\text{nm}}$ / nm	Green Phosphor $(\text{Lu,Ce})_3(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	Red Phosphor $(\text{Ca,Sr,Ba})_2\text{Si}_2\text{N}_6$ (10% Ca, 47% Sr, 50% Ba, 3% Eu)	$\lambda_{\text{dom}}^{\text{red}}$ / nm (red phosphor concentration)	green phosphor concentration / red	CIE x	CIE y	CCT / K	CRI	rel. eff.
Comparative example 1	446	$(\text{Lu,Ce})_3(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$(\text{Ca,Sr,Ba})_2\text{Si}_2\text{N}_6$ (10% Ca, 47% Sr, 50% Ba, 3% Eu)	601.5	15.5%	4.45:1	0.461	0.408	2669	79 100.0%
Comparative example 3	446	$(\text{Lu,Ce})_3(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$(\text{Sr,Ca})\text{AlSiN}_3$ (80% Sr, 0.5% Eu)	599.4	18.0%	2.5:1	0.459	0.416	2754	80 101.4%
Comparative example 4	446	$(\text{Lu,Ce})_3(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$(\text{Sr,Ca})\text{AlSiN}_3$ (80% Sr, 0.5% Eu)	600.3	16.8%	2.6:1	0.461	0.408	2655	81 101.5%
Inventive example 2	446	$(\text{Lu,Ce})_3(\text{Al,Ga})_5\text{O}_{12}$ (25% Ga, 2.5% Ce)	$\text{Sr}(\text{Sr}_{0.8}\text{Ca}_{0.2})\text{Si}_2\text{Al}_2\text{N}_6\text{Eu}$ (1.2% Eu)	602.4	16.5%	5.5:1	0.462	0.408	2649	82 105.8%

Fig. 8



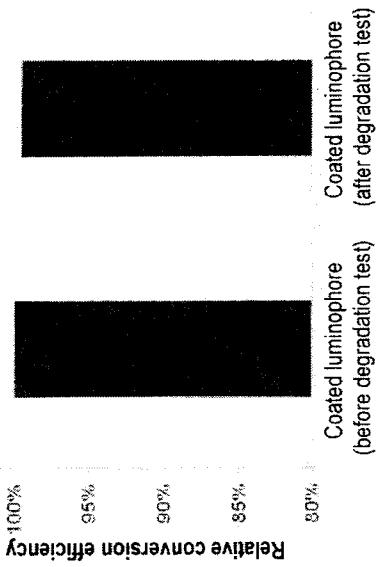
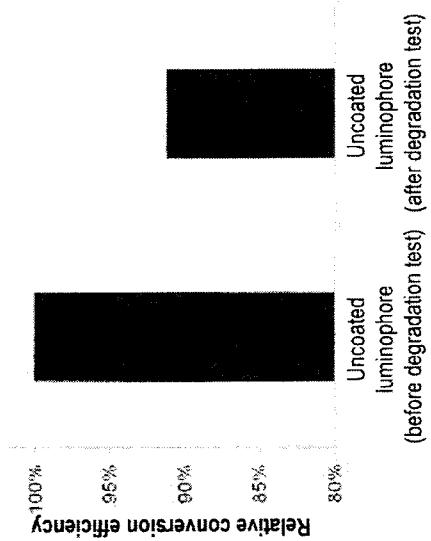
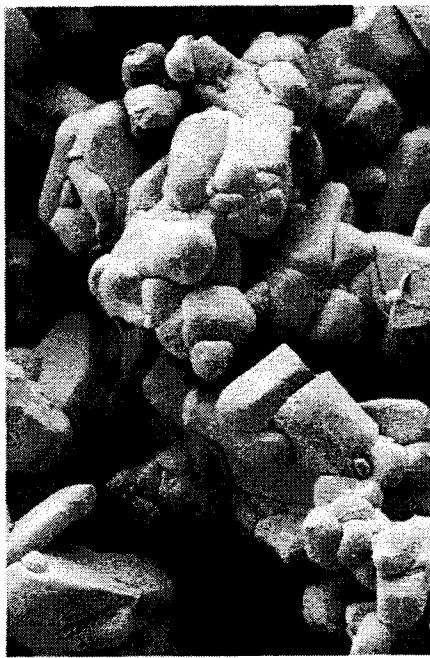


Fig. 9



Uncoated
Before degradation test

Fig. 10a

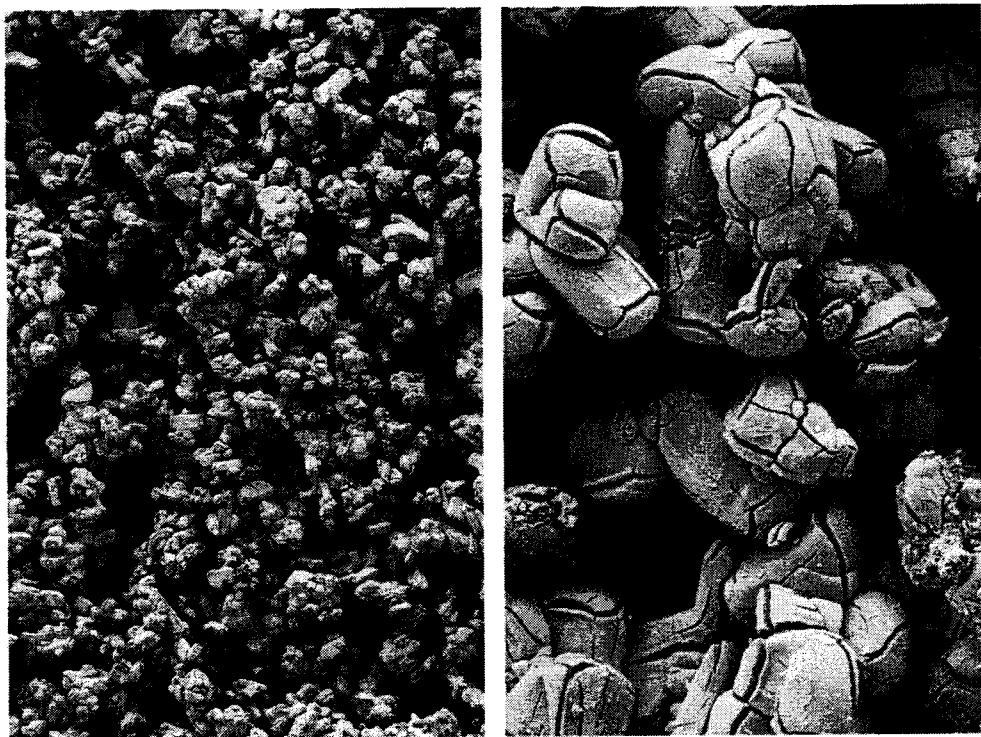


Fig. 10b



Fig. 11a

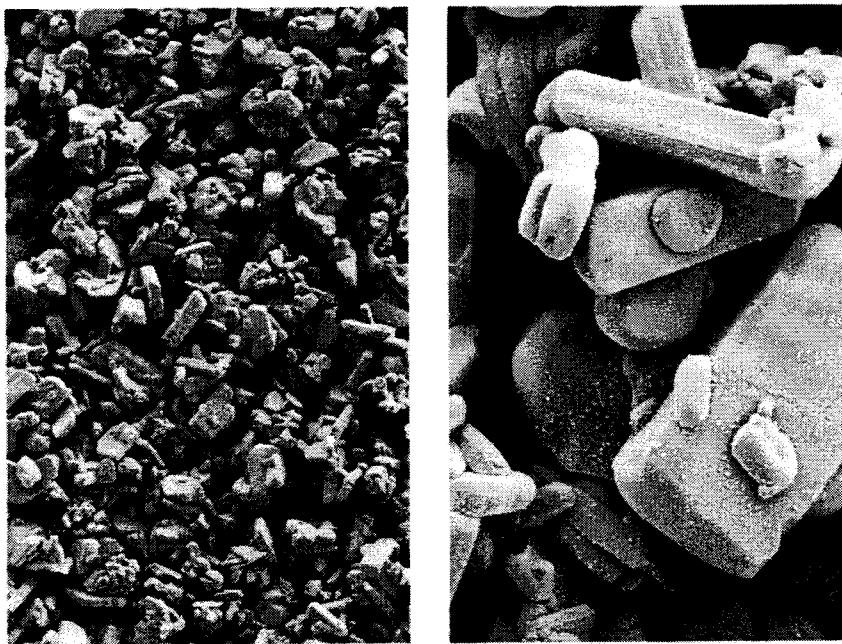


Fig. 11b

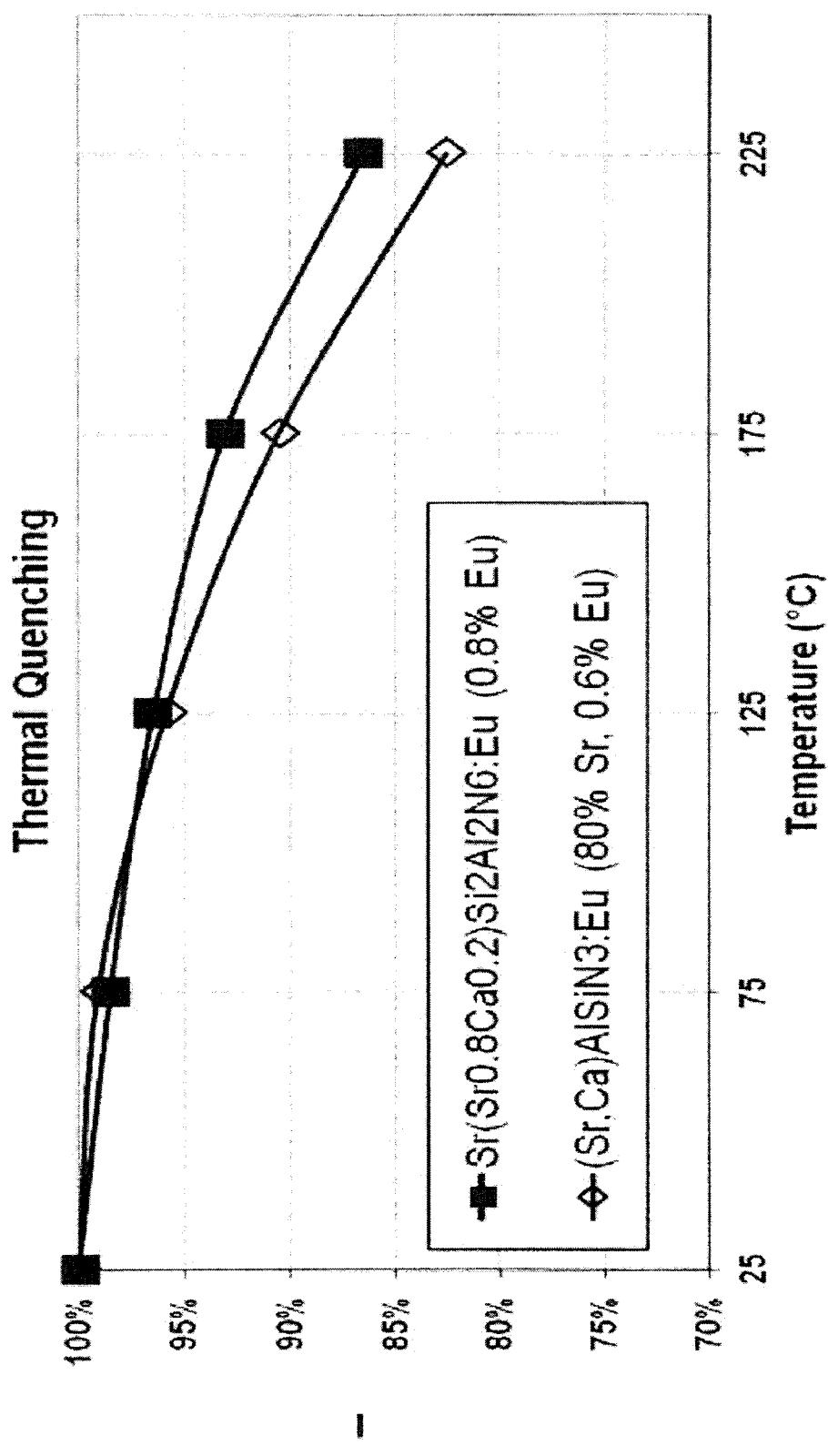
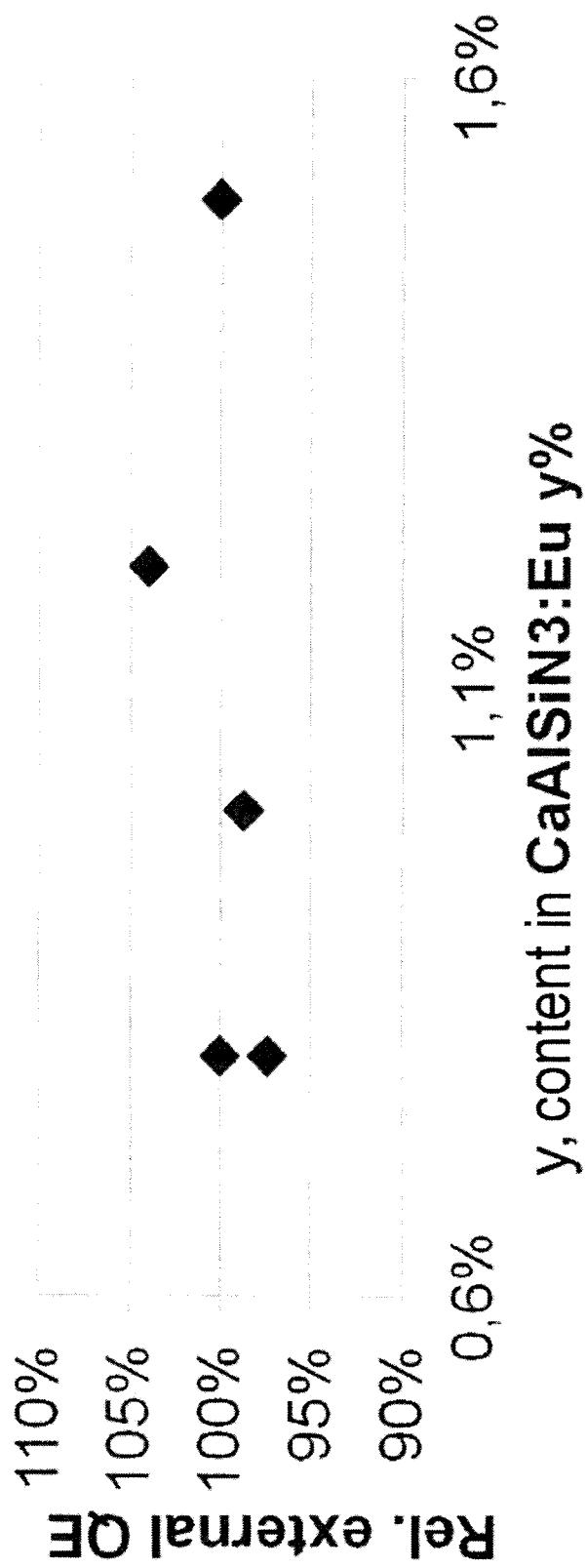


Fig. 12

Fig. 13



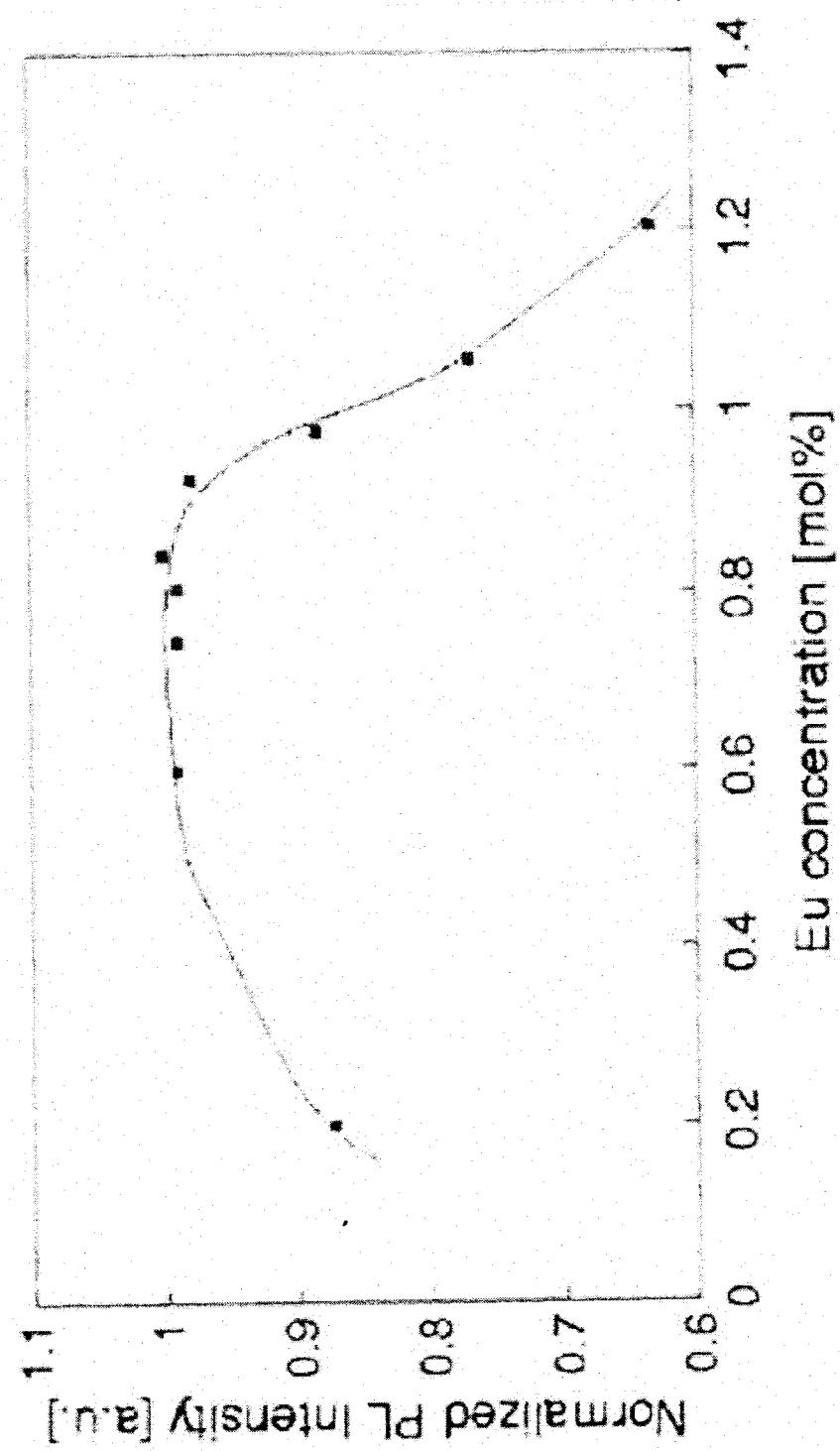


Fig. 14

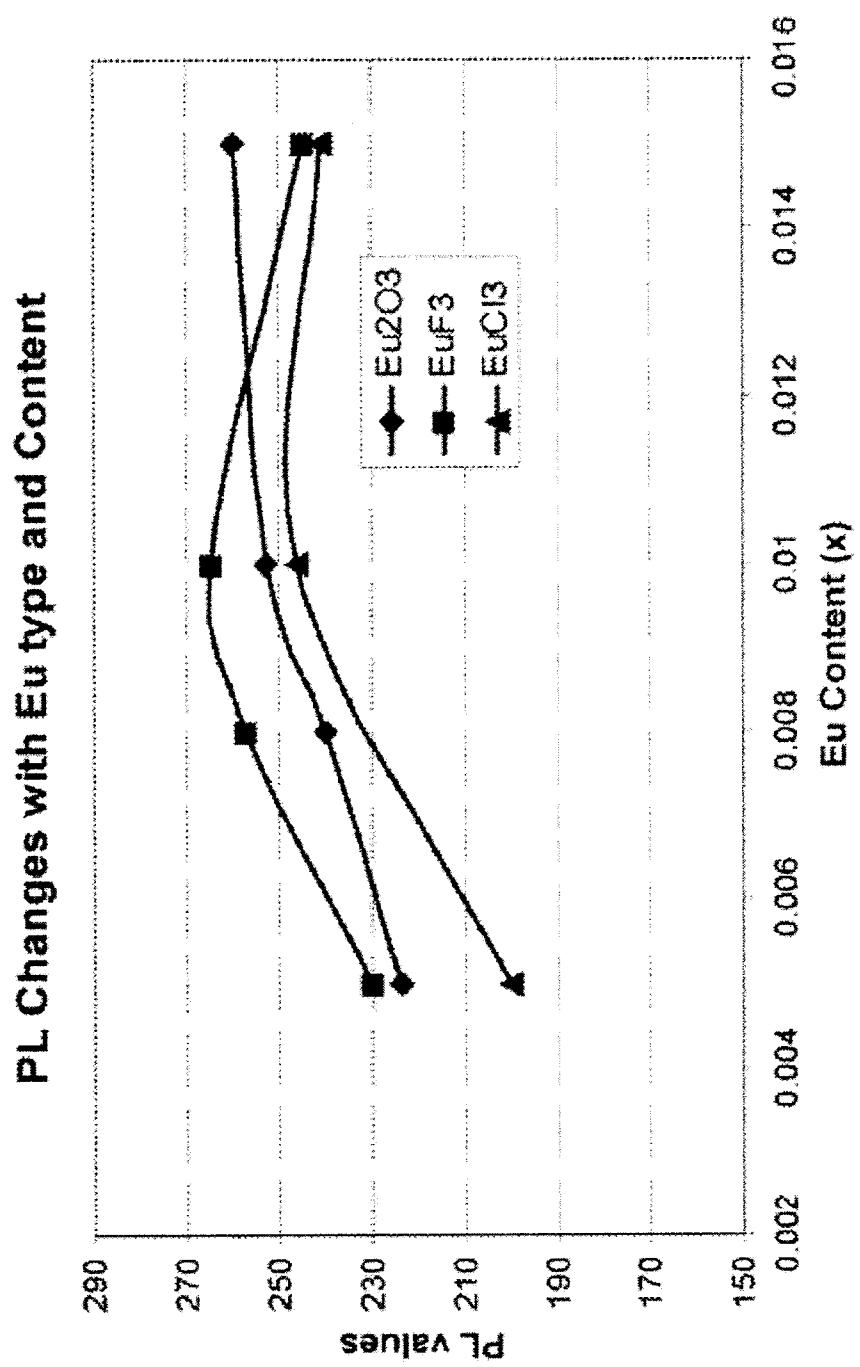


Fig. 15

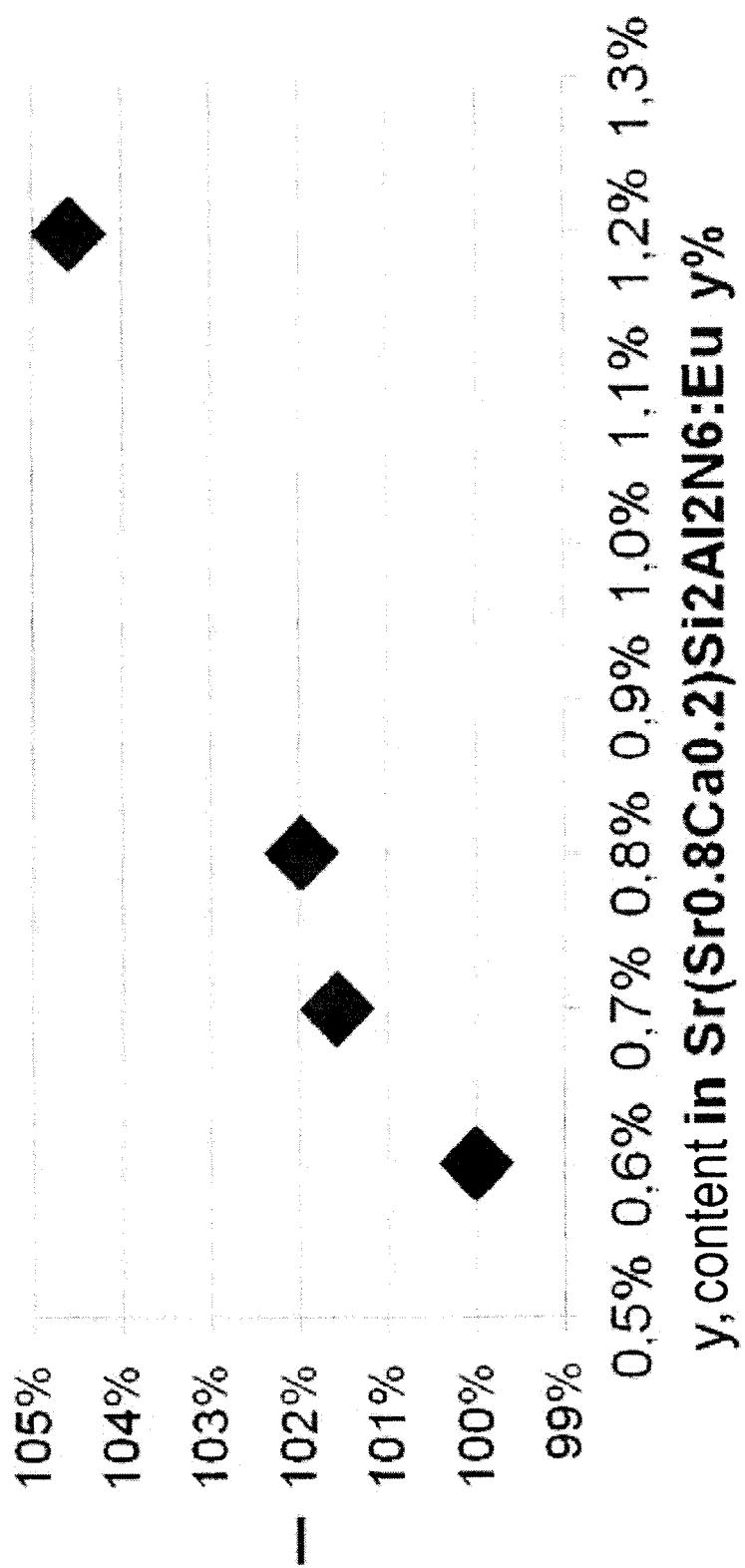
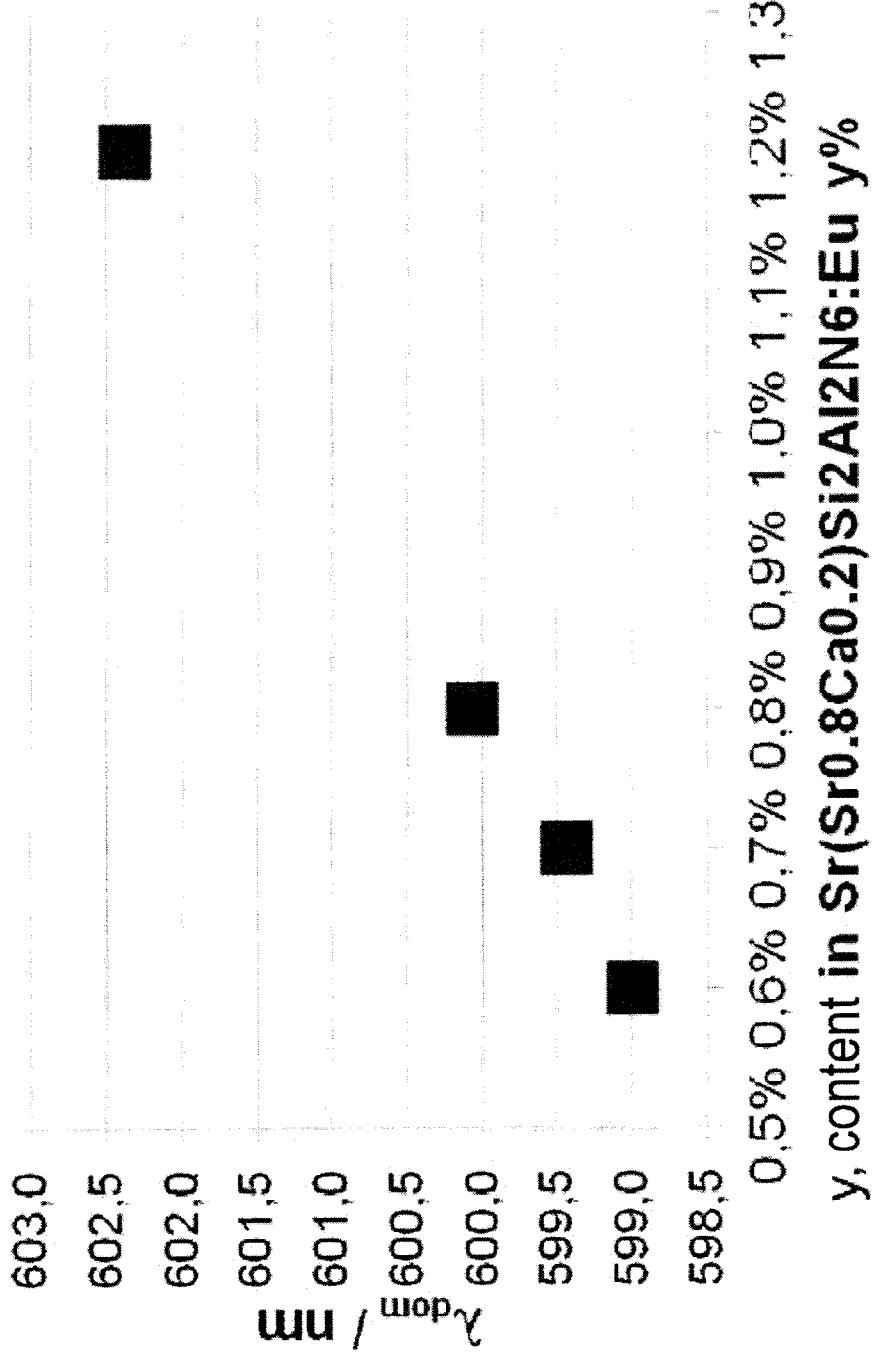


Fig. 16

Fig. 17



0,5% 0,6% 0,7% 0,8% 0,9% 1,0% 1,1% 1,2% 1,3%
 y, content in **Sr_{0.8}Ca_{0.2}Si₂Al₂N₆:Eu** y%

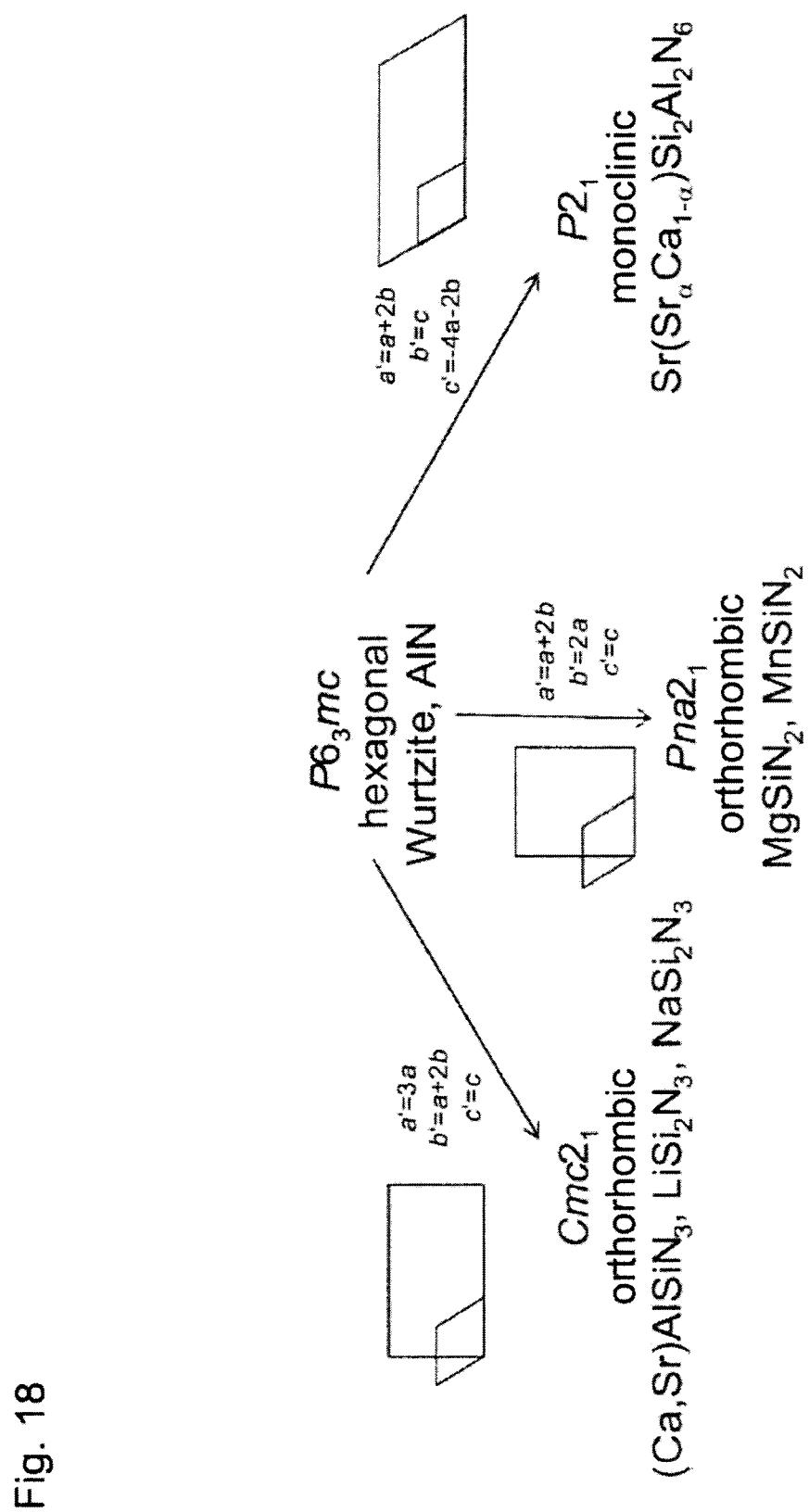


Fig. 19

Structural data from the ICSD database; some compounds in the $Cmc\bar{2}1$ space group, all in the NaSiO_3 structure type

Formula	a (Å)	b (Å)	c (Å)	Bibliographic data
$\text{Ca}(\text{Al}_{0.55}\text{Si}_{1.34})\text{N}_3$	9.7558	5.6473	5.0524	Chem. Mater. (1,1989 - (2008) 20, p6704 -p6714)
$(\text{Ca}_{0.8}\text{Ce}_{0.1}\text{Li}_{0.1})(\text{Al}_{0.5}\text{Si}_{1.3})\text{N}_3$	9.795	5.6465	5.0590	Chem. Mater. (1,1989 - (2008) 20, p6704 -p6714)
$(\text{Ca}_{0.8}\text{Ce}_{0.1}\text{Li}_{0.1})(\text{AlSi})\text{N}_3$	9.813	5.6667	5.08378	Journal of Alloys Compd. (2009), 475, p434 -439
$(\text{Sr}_{0.2}\text{Ca}_{0.8})(\text{AlSi})\text{N}_3$	9.8297	5.6879	5.1085	Journal of Alloys Compd. (2009), 475, p434 -439
$(\text{Sr}_{0.8}\text{Ca}_{0.2})(\text{AlSi})\text{N}_3$	9.8152	5.73644	5.14905	Journal of Alloys Compd. (2009), 475, p434 -439
$\text{Li}_{0.34}\text{Ca}_{0.66}\text{Eu}_{0.01}\text{Al}_{0.01}\text{Si}_{1.38}\text{N}_3$	9.2204	5.3107	4.7861	Journal of Sol. State Chem. (2009), 182 (2), p301 - 311
$\text{Sr}_{0.22}\text{Eu}_{0.78}(\text{AlSi})\text{N}_3$	9.8087	5.756	5.16614	Journal of Alloys Compd. (2009), 475, p434 -439
$(\text{Sr}_{0.99}\text{Eu}_{0.01})(\text{AlSi})\text{N}_3$	9.843	5.7603	5.177	Journal of Sol. State Chem. (2009), 181 (8), p1848 - 1852

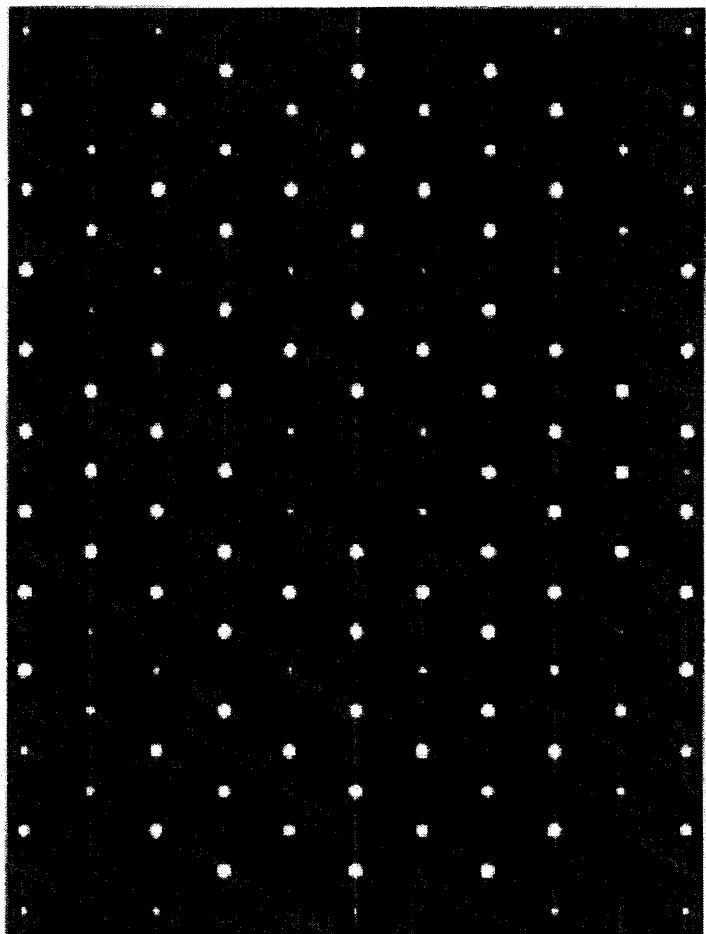


Fig. 20a

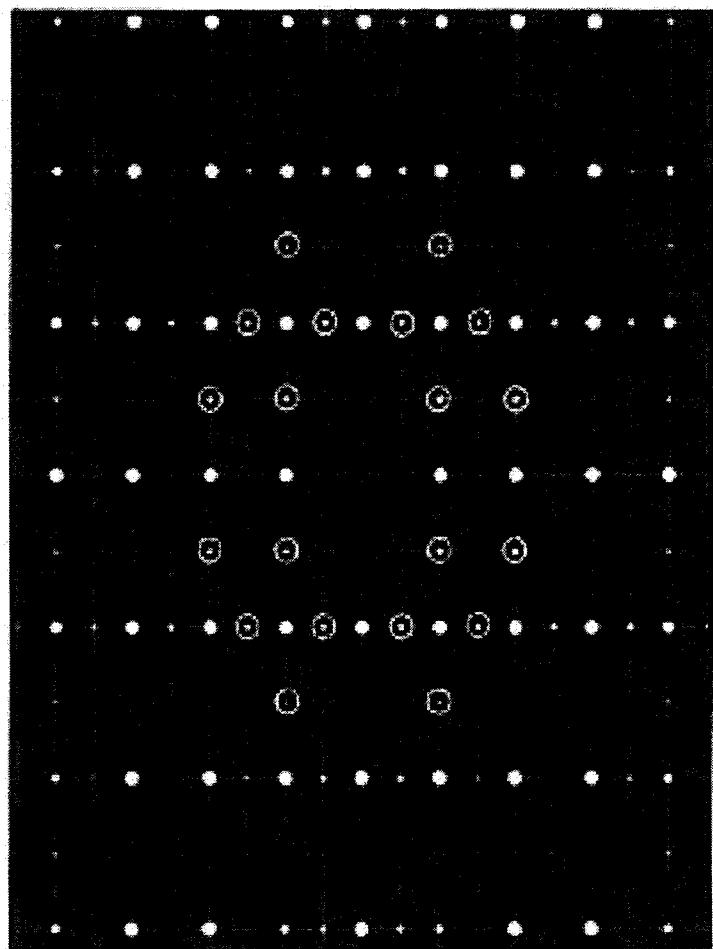


Fig. 20b

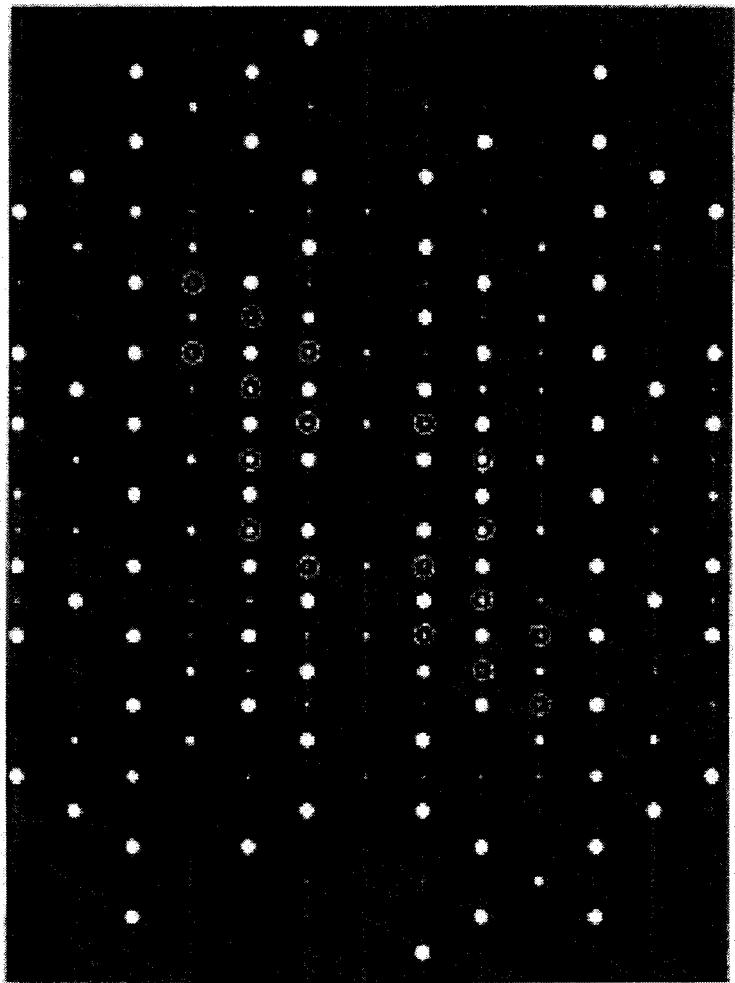


Fig. 20c

Single crystal diffractometry solution of the novel structure $\text{Sr}(\text{Sr}_\alpha \text{Ca}_{1-\alpha})\text{Si}_2\text{Al}_2\text{N}_6$

Details of the x-ray crystallography solution	
Lattice parameters	
<i>a</i>	574.93(3) pm
<i>b</i>	516.35(3) pm
<i>c</i>	1136.1(2) pm
β	120.341(2) °
Formula units <i>Z</i>	2
Radiation	Mo K _a , $\lambda = 71.073$ pm
Measured/independent reflections	8331/ 2053
Measured section of reciprocal space	-8 < <i>h</i> < 8; -7 < <i>k</i> < 8; -16 < <i>l</i> < 17
$R_{\text{obs}}/wR_{\text{obs}}$	6.46%/11.38%
GoF	1.97

Fig. 21

Distances in the novel structure
 $\text{Sr}(\text{Sr}_\alpha \text{Ca}_{1-\alpha})\text{Si}_2\text{Al}_2\text{N}_6$

Important interatomic distances in the novel structure $\text{Sr}(\text{Sr}_\alpha \text{Ca}_{1-\alpha})\text{Si}_2\text{Al}_2\text{N}_6$								
Si1/Al1-	N1	175 pm	Si2/Al2-	N1	202 pm	Si3/Al3-		
N2	181 pm	N2	176 pm	N3	194 pm	N1	169 pm	
N3	188 pm	N4	178 pm	N4	183 pm			
N5	182 pm	N6	171 pm	N5	167 pm			
Si4/Al4-	N2	187 pm	Si1	N1	259 pm	Si2/Ca2	N2	250 pm
N3	159 pm	N3	263 pm	N4	281 pm			
				N5_1	264 pm	N6_1	253 pm	
				N5_2	270 pm	N6_2	272 pm	
				N5_3	277 pm	N6_3	263 pm	
				N5_4	297 pm			

Fig. 22

Fig. 23

Crystallographic parameters (single crystallography)

Crystallographic data of $\text{Sr}_2\text{Si}_2\text{Al}_2\text{N}_4$

Space group: $P2_1$
Crystal system: monoclinic
Structure type: $\text{Sr}_2(\text{Sr}_2\text{Al}_2\text{N}_4)\text{Si}_2\text{Al}_2\text{N}_4$
Lattice parameters:
 $a = 574.93 \text{ pm}$
 $b = 516.35 \text{ pm}$
 $c = 1136.1 \text{ pm}$
 $\beta = 120.341^\circ$

	$A/\text{Å}^3$	$B/\text{Å}^2 \cdot 10^4$	$C/\text{Å}^2 \cdot 10^4$
Al_1	0, 5000	0, 0349	0, 9764
Al_2	0, 5000	0, 0324	0, 9720
Al_3	0, 5000	0, 0773	0, 9255
Al_4	0, 5000	0, 0615	0, 0037
Si_1	0, 5000	0, 0349	0, 9764
Si_2	0, 5000	0, 0324	0, 9720
Si_3	0, 5000	0, 6773	0, 9255
Si_4	0, 5000	0, 5000	0, 0037
N_1	0, 5000	0, 0000	0, 9950
N_2	0, 5000	0, 0000	0, 9950
N_3	0, 5000	0, 0000	0, 9950
N_4	0, 5000	0, 0000	0, 9950
N_5	0, 5000	0, 0000	0, 9950
N_6	0, 5000	0, 0000	0, 9950
Sr_1	2.3	1, 0000	0, 9806
Sr_2	2.3	0, 6249	0, 6394
O_1	2.3	0, 3750	0, 6894

Crystallographic data of $\text{Sr}_2\text{Si}_2\text{Al}_2\text{N}_4$, known from the literature

(Structure database ICSD 41-9410)
Space group: $Cm2_1$
Crystal system: orthorhombic
Structure type: Na_2SiO_3
Lattice parameters:
 $a = 984.3 \text{ pm}$
 $b = 576.03 \text{ pm}$
 $c = 517.7 \text{ pm}$

	$A/\text{Å}^3$	$B/\text{Å}^2 \cdot 10^4$	$C/\text{Å}^2 \cdot 10^4$
Al_1	0, 5000	0, 1727	0, 1518
Al_2	0, 5000	0, 1777	0, 1518
N_1	0, 5000	0, 0000	0, 2919
N_2	0, 0000	0, 2714	0, 1177
N_3	0, 0000	0, 2714	0, 1713
N_4	0, 0000	0, 1897	0, 0000
Sr_1	0, 9900	0, 0000	0, 829
Sr_2	0, 0100	0, 0000	0, 5000

Crystallographic data/location parameters of the novel phosphor of the invention (left) - and as a comparison the crystallographic data/location parameters of the phosphor SrAlSiN_3 , known from the literature (on the right).

Fig. 24a

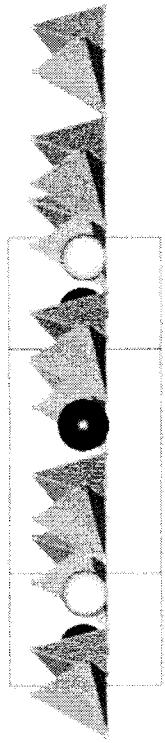


Fig. a) Phosphor of the invention
 $Sr(Sr_{\alpha}Ca_{1-\alpha})Si_2Al_2N_6:Eu$, view of the layers
a → c
b

Fig. 24b

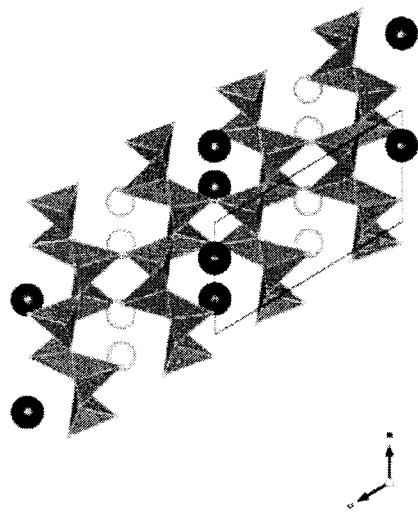


Fig. b) Phosphor of the
invention $Sr(Sr_{\alpha}Ca_{1-\alpha})Si_2Al_2N_6:Eu$, [010] viewing
direction

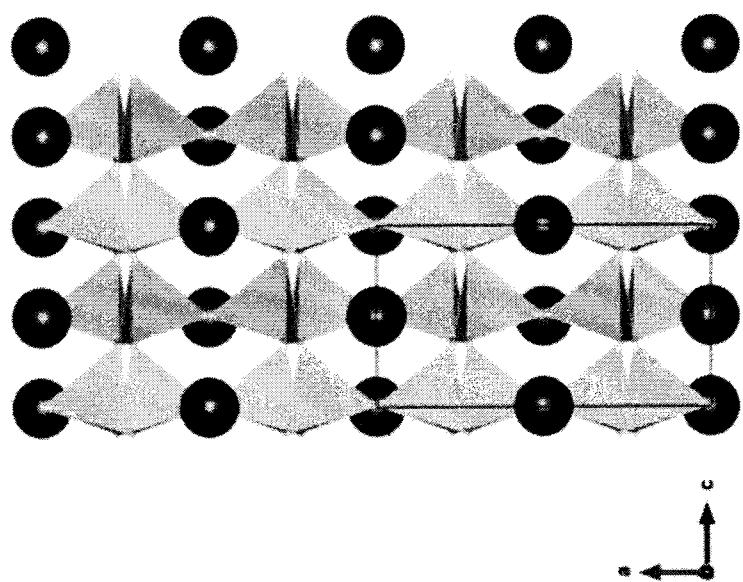


Fig. c) literature-known
phosphor SCASN

Fig. 24c

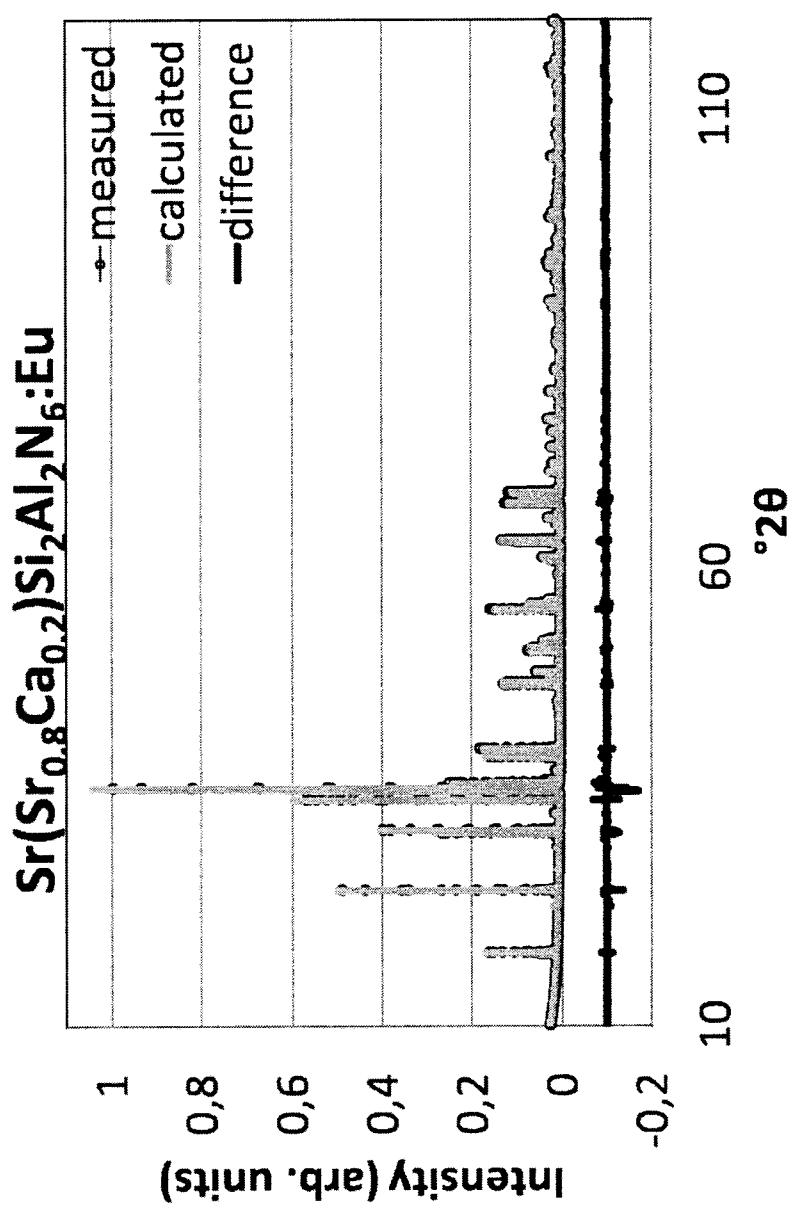


Fig. 25

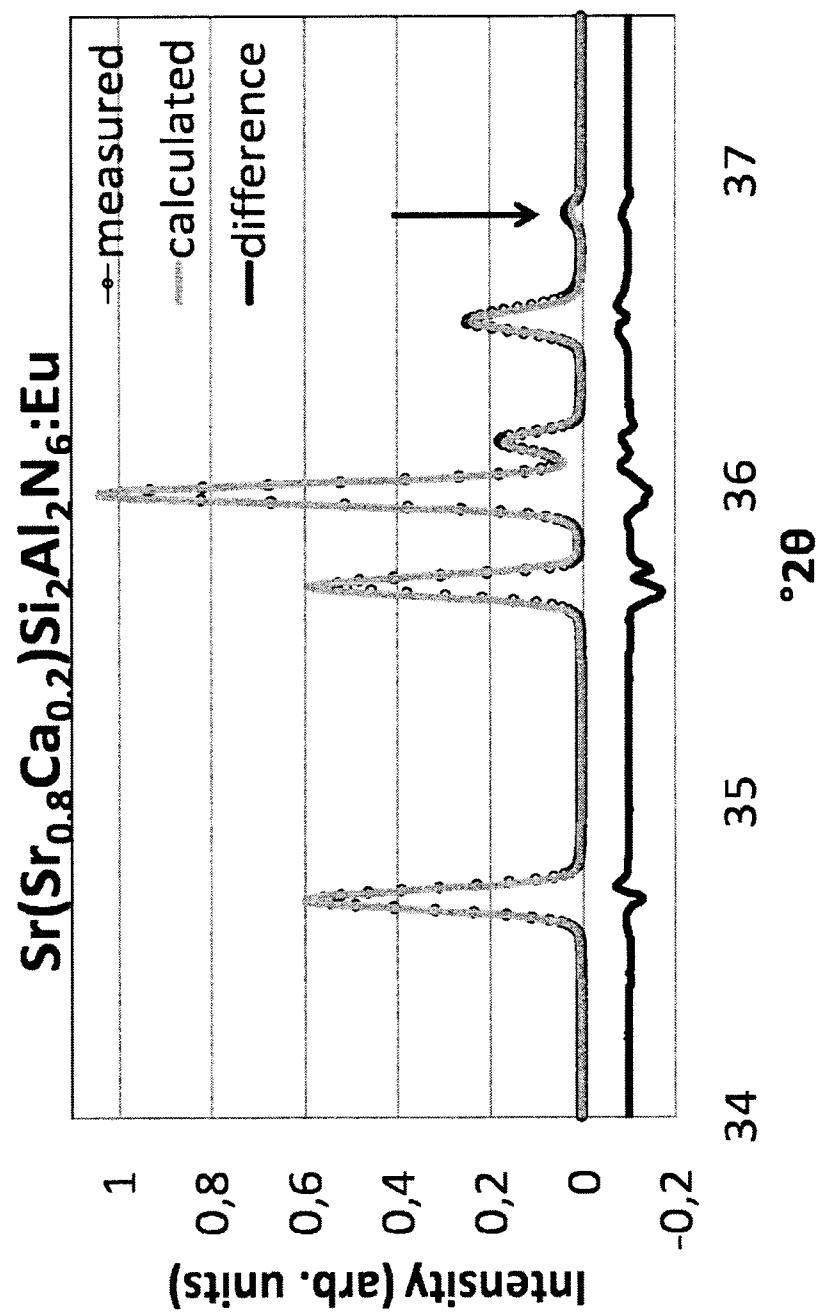


Fig. 26

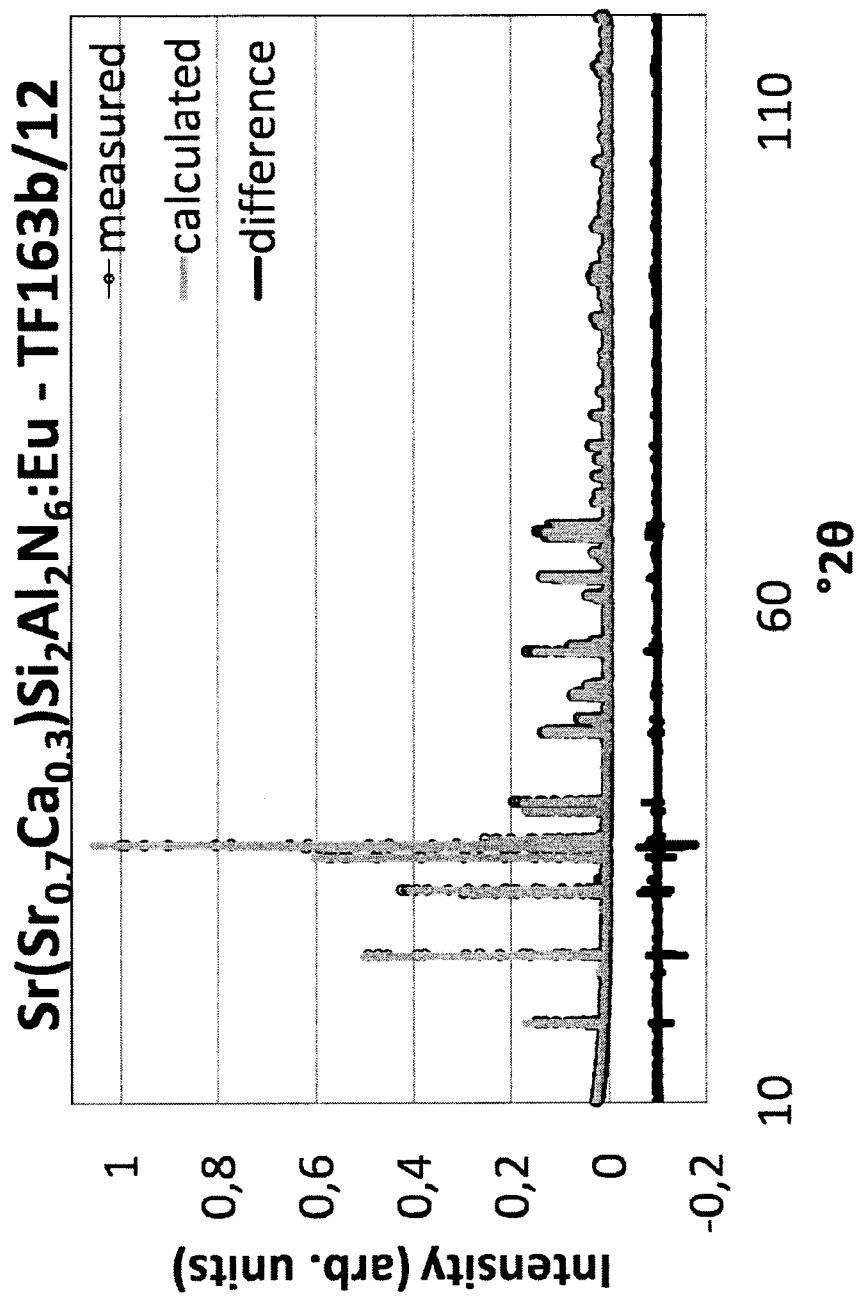


Fig. 27

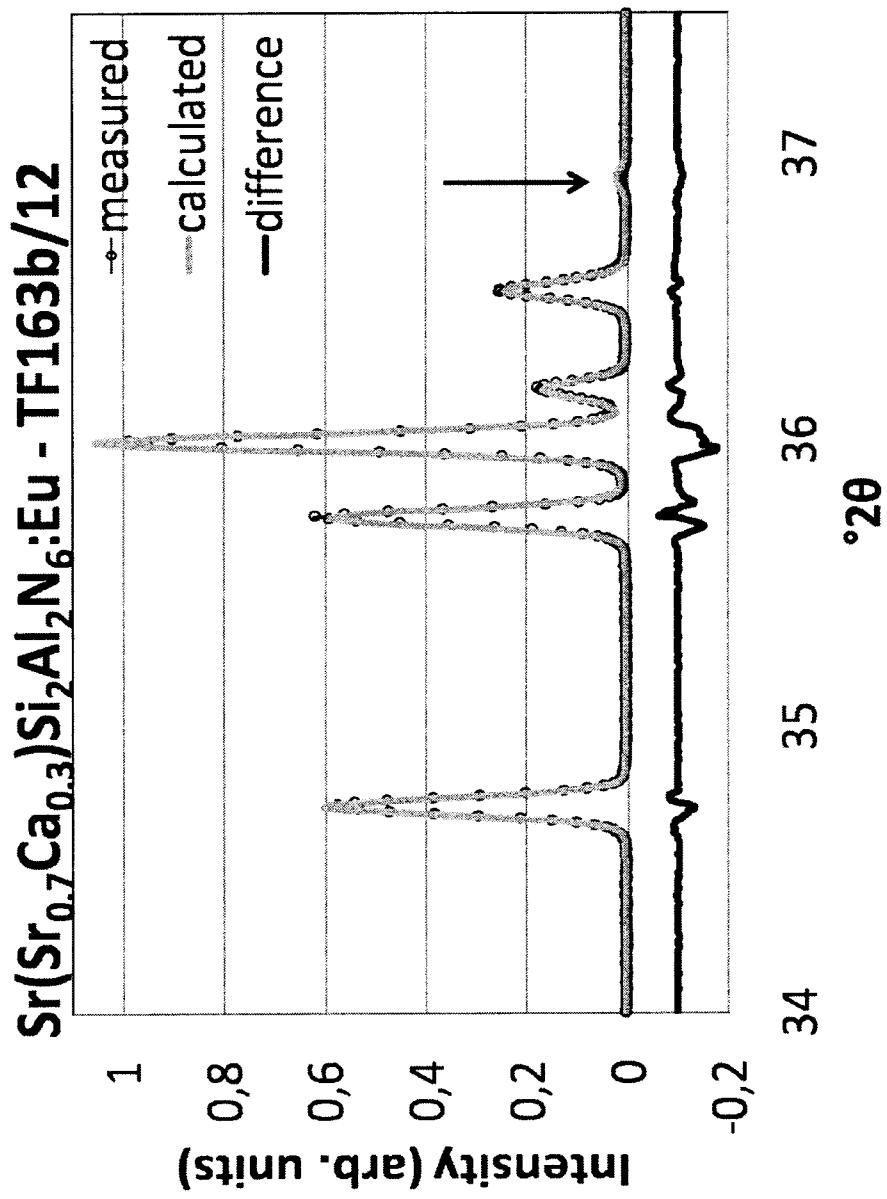


Fig. 28

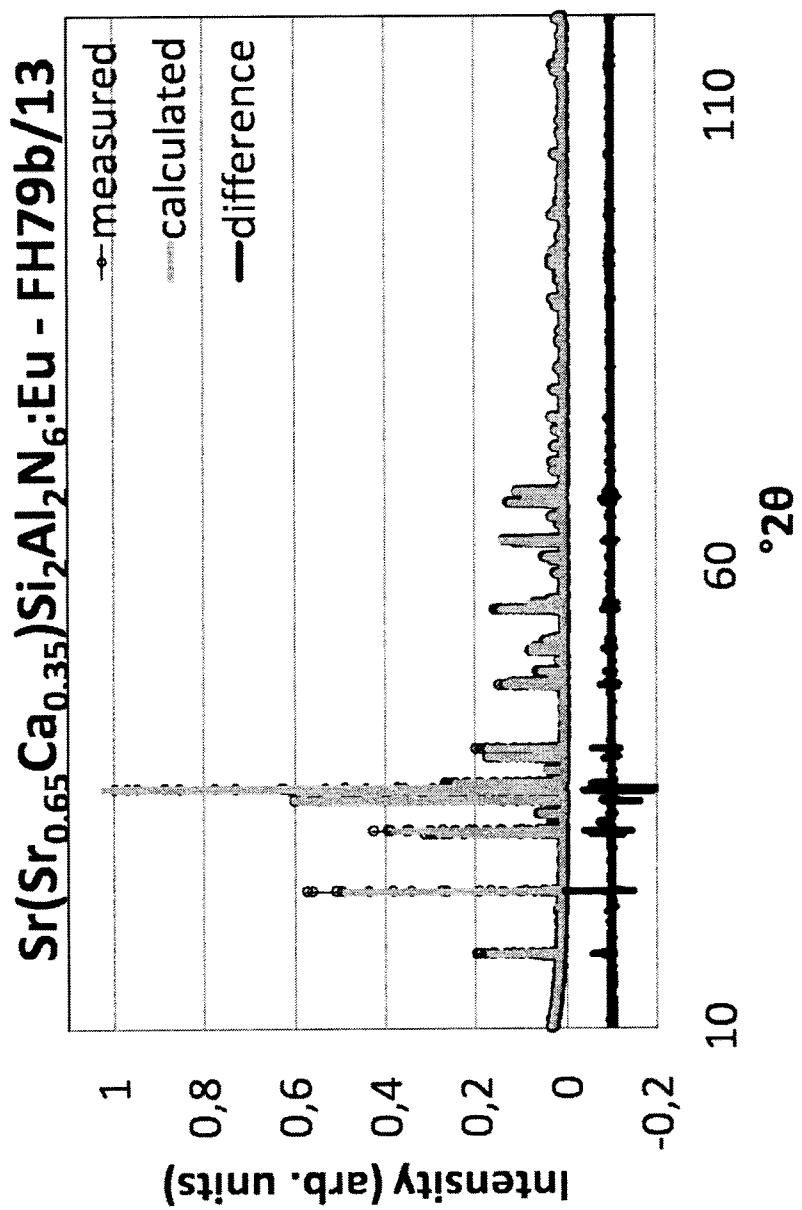


Fig. 29

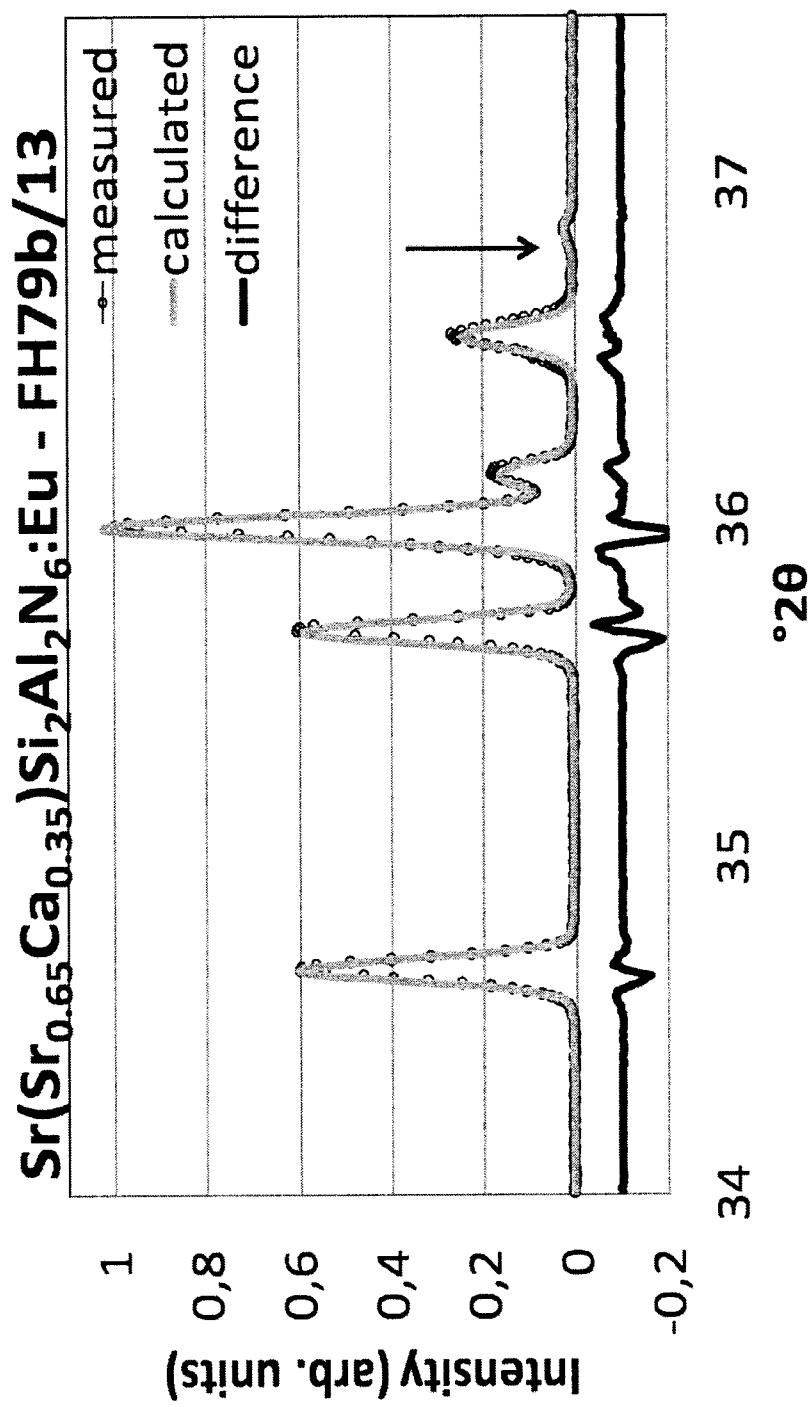


Fig. 30

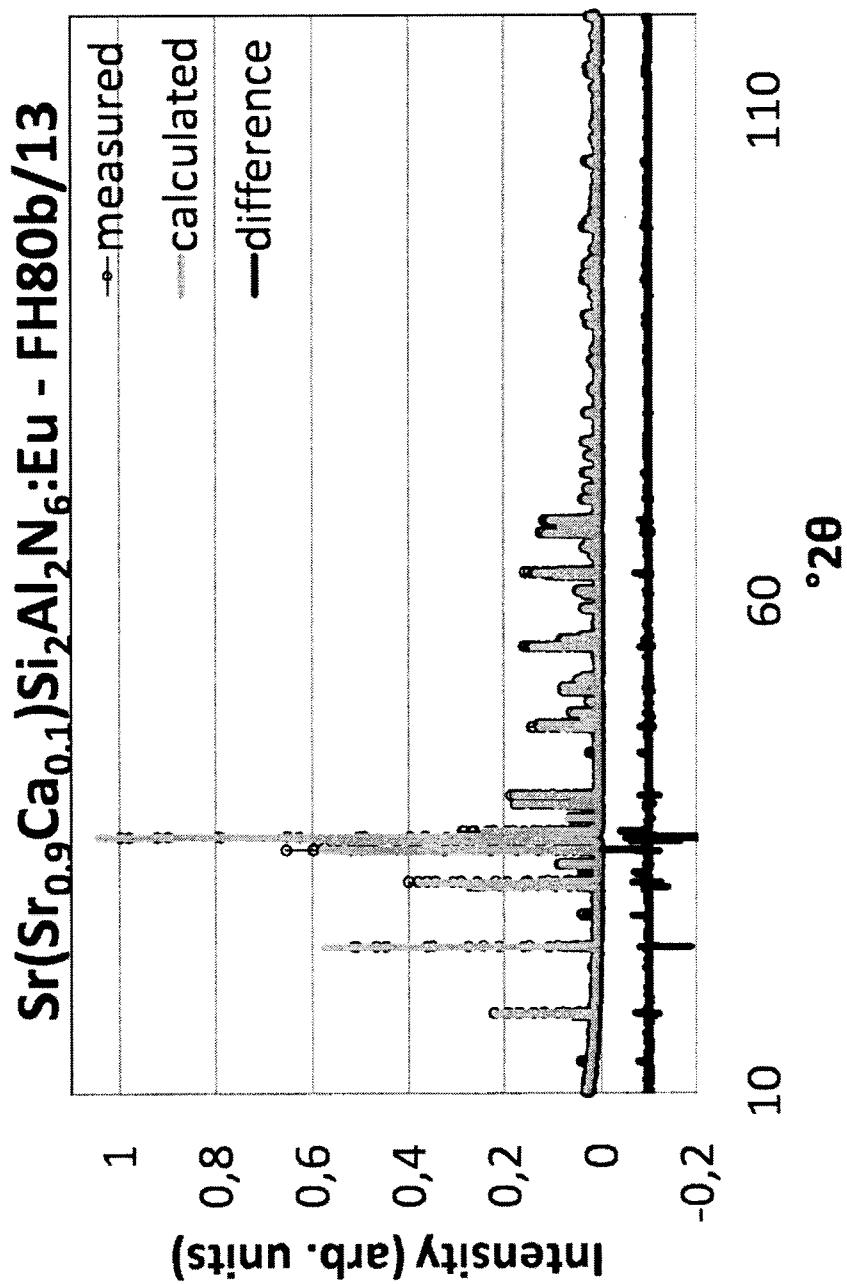


Fig.31

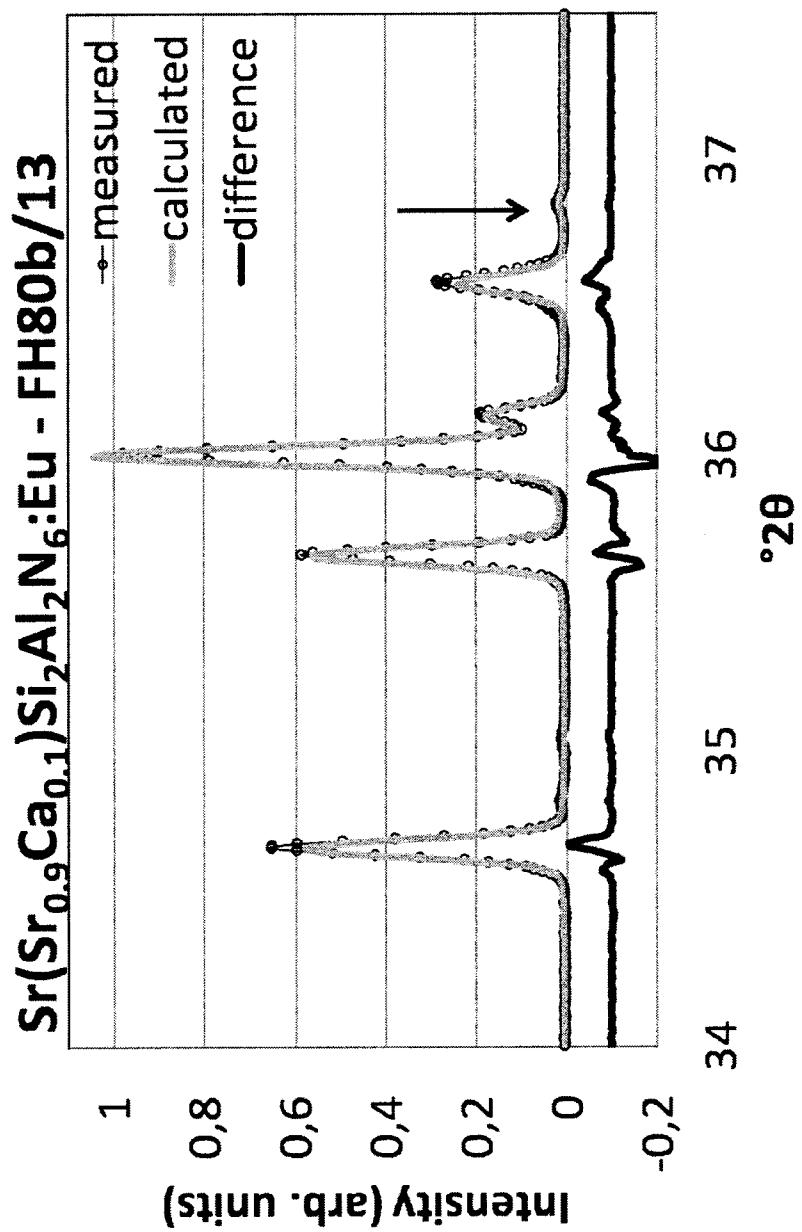


Fig.32

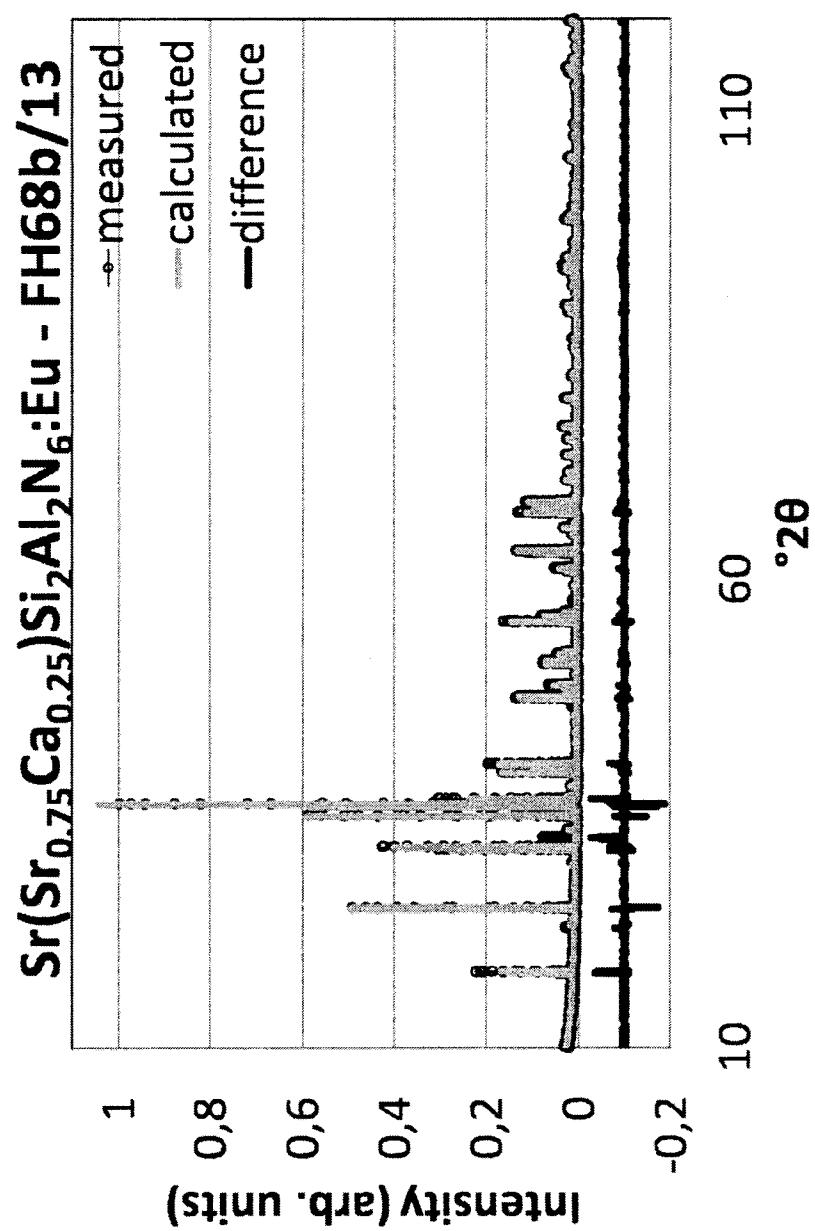


Fig.33

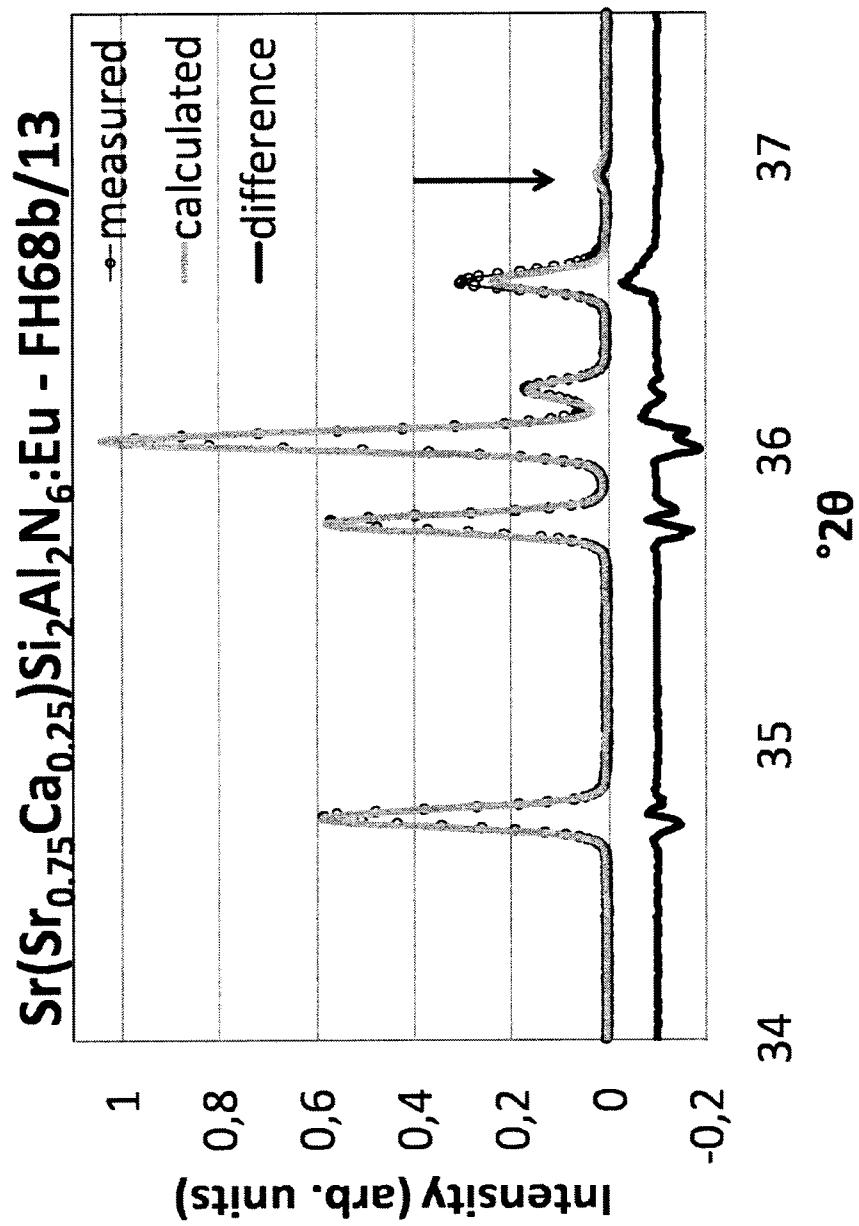


Fig.34

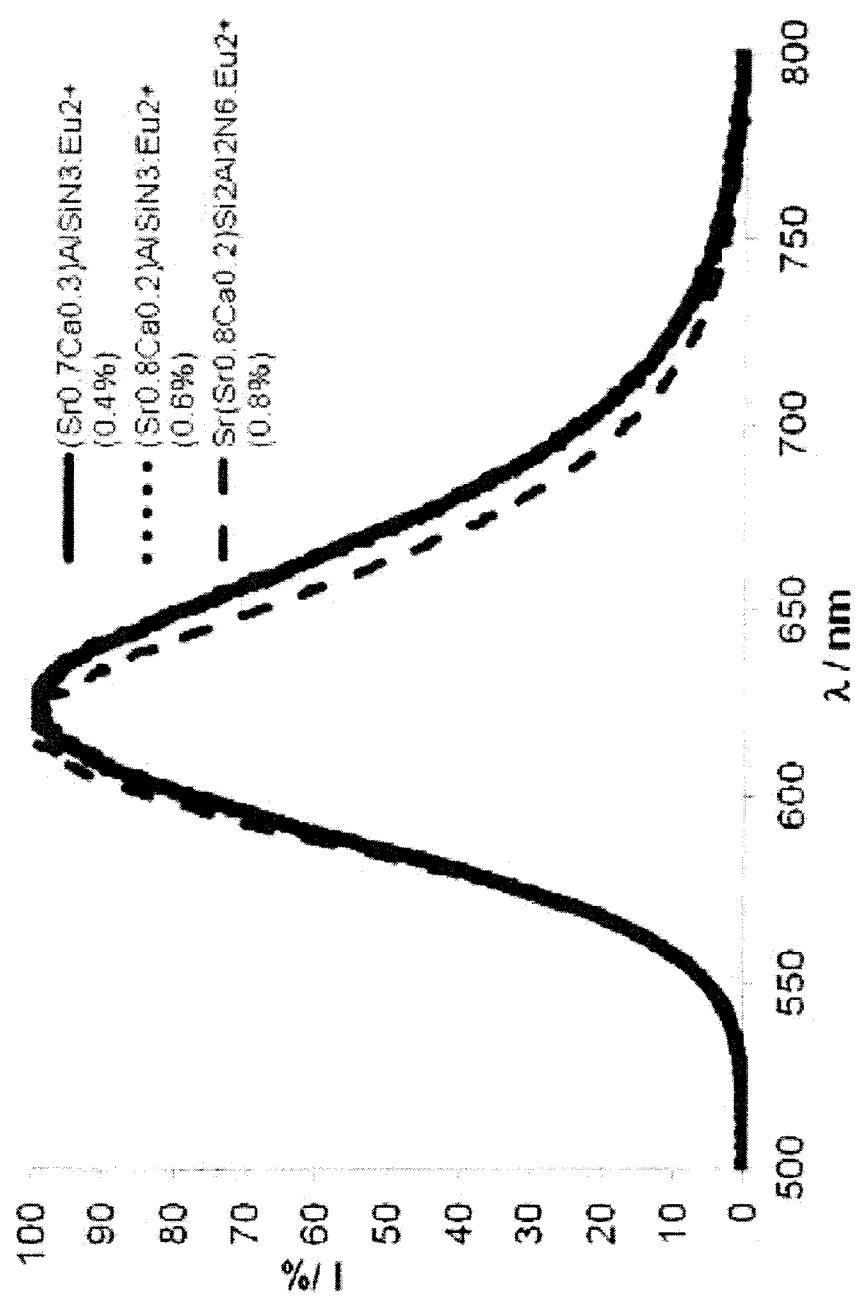


Fig. 35a

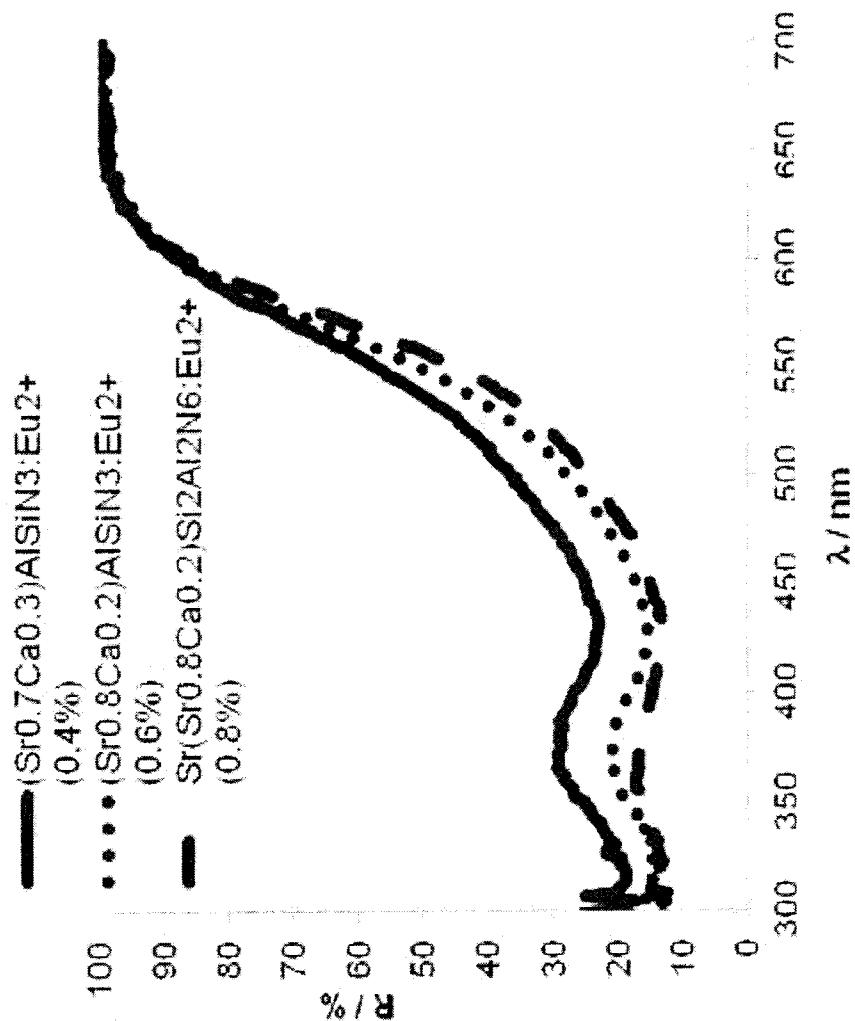


Fig. 35b

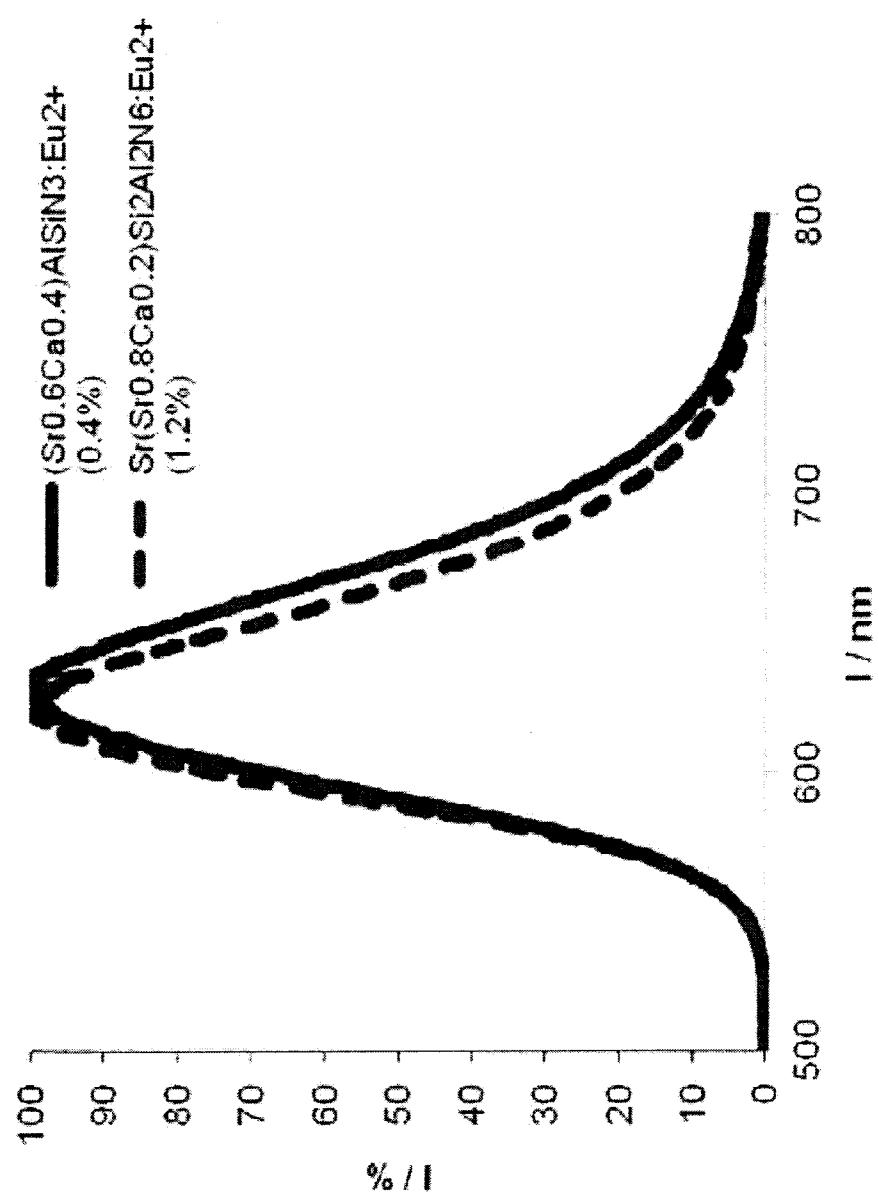


Fig. 36a

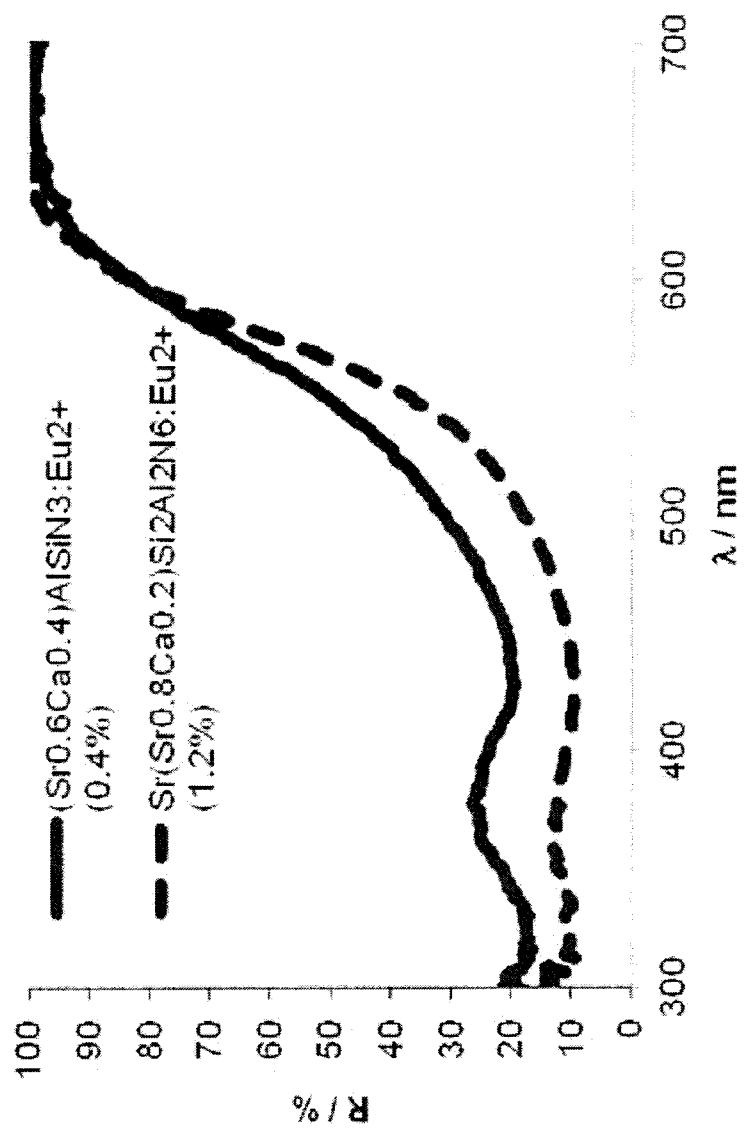


Fig. 36b

FIG 36b

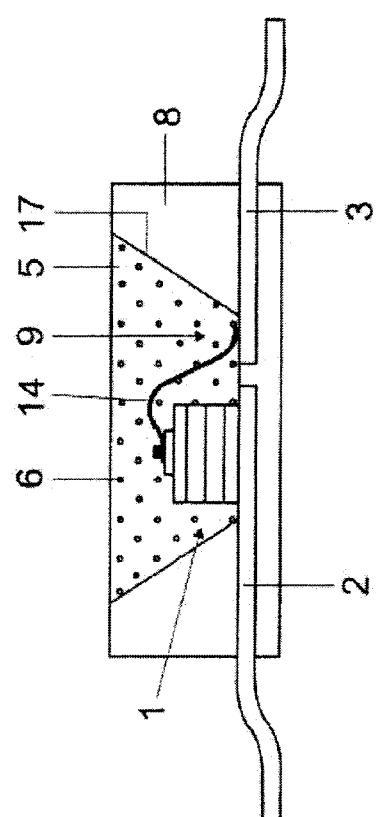


Fig. 37

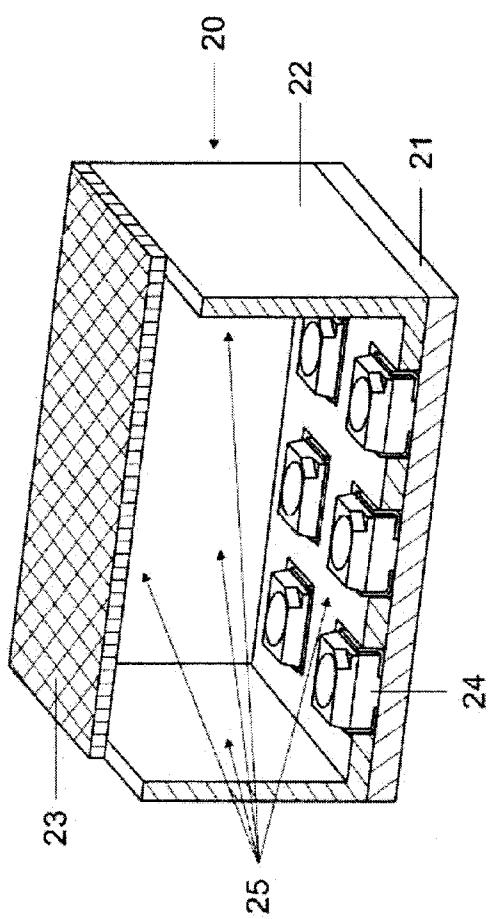


Fig. 38

Fig. 39a/1

Experiment	y	Composition		Weights ^g										Optical characterization			
		Ca _{1-x} AlSi _y Eu	or	Sr(Sr _x Ca _{1-x})Si ₂ Al ₂ Ne _y Eu	(a=2x-1)	AlF3	SrF2	BaF2	LiBF4	Li2B4O7	Si3N4	AlN	Eu2O3	CIE x	CIE y		
FH 005_13	0,6%	0,8	x	Ca3N2	2,746	21,552	0,875	0,436	0	0	12,993	11,104	0,293	50	0,632	0,365	
FH 006_13	0,6%	0,8			2,730	21,428	0,290	1,301	0	0	12,918	11,040	0,292	50	0,632	0,365	
FH 007_13	0,6%	0,8			2,729	21,416	0,232	1,387	0	0	12,911	11,034	0,292	50	0,634	0,364	
FH 008_13	0,6%	0,8			2,748	21,564	0,934	0,349	0	0	13,000	11,111	0,294	50	0,632	0,365	
FH 009_13	0,6%	0,8			2,738	21,492	1,163	0	0	0	12,957	11,357	0,293	50	0,631	0,366	
FH 010_13	0,6%	0,8			2,646	20,769	1,124	1,682	0	0	12,521	10,975	0,283	50	0,631	0,366	
FH 011_13	0,6%	0,8			2,574	20,202	1,641	2,454	0	0	12,179	10,676	0,275	50	0,633	0,364	
FH 012_13	0,6%	0,8			2,506	19,665	2,129	3,185	0	0	11,855	10,392	0,268	50	0,632	0,364	
FH 013_13	0,6%	0,8			2,731	21,434	0	0	0	0	1,295	12,922	11,326	0,292	50	0,635	0,363
FH 014_13	0,6%	0,8			2,735	21,463	0,581	0	0	0	0,648	12,939	11,342	0,292	50	0,633	0,364
FH 015_13	0,6%	0,8			2,719	21,340	0,000	0,864	0	0	0,645	12,865	11,277	0,290	50	0,634	0,364
FH 016_13	0,6%	0,8			2,725	21,389	0,387	0,578	0	0	0,432	12,895	11,303	0,291	50	0,633	0,364
FH 029_13	0,6%	0,8			2,671	20,961	0	0	2,369	0	0	12,637	11,077	0,285	50	0,633	0,364
FH 030_13	0,6%	0,8			2,704	21,223	0,574	0	1,199	0	0	12,795	11,215	0,289	50	0,631	0,366
FH 031_13	0,6%	0,8			2,689	21,103	0,000	0,854	1,193	0	0	12,722	11,152	0,287	50	0,630	0,367
FH 032_13	0,6%	0,8			2,701	21,195	0,000	0	1,198	0	0,640	12,778	11,200	0,289	50	0,636	0,361
FH 037_13	0,4%	0,6			5,741	16,896	0,305	1,368	0	0	0	13,581	11,905	0,204	50	0,636	0,360
FH 038_13	0,4%	0,6			5,738	16,886	0,244	1,458	0	0	0	13,573	11,897	0,204	50	0,635	0,361
FH 039_13	0,4%	0,6			5,776	16,999	0,920	0,459	0	0	0	13,664	11,977	0,206	50	0,637	0,359
FH 040_13	0,4%	0,6			5,779	17,009	0,982	0,367	0	0	0	13,672	11,984	0,206	50	0,638	0,358
FH 041_13	0,4%	0,6			5,793	17,048	0,000	0	0	0	1,239	13,703	12,011	0,206	50	0,636	0,359
FH 042_13	0,4%	0,6			5,793	17,049	0,615	0	0	0	0,620	13,704	12,012	0,206	50	0,640	0,356
FH 043_13	0,4%	0,6			5,758	16,946	0	0,915	0	0	0,616	13,621	11,940	0,205	50	0,639	0,357
FH 044_13	0,4%	0,6			5,785	17,025	0	0	0	0,619	0,686	13,685	11,995	0,206	50	0,642	0,354

Fig. 39a/2

Experiment	y	Composition										Optical characterization						
		Ca _{3-x} SrAlSiN _{3-y} Eu	of	Si _x (Sr _{1-x})Si ₂ Al ₂ N _{6-y} Eu	($\alpha=2x-1$)	Ca3N2	Si3N2	AlF3	SiF2	BaF2	Li2B4O7	LiBF4	Si3N4	AlN	Eu2O3	Total	CIE x	CIE y
FH 046_13	0.4%	x	0.6	5.747	16.913	0	1.205	0	0.418	0	13.595	11.917	0.205	50	0.639	0.358	0.358	0.358
FH 047_13	0.4%	0.6	5.737	16.885	0	1.458	0	0.245	0	13.573	11.897	0.204	50	0.639	0.358	0.358	0.358	
FH 048_13	0.4%	0.6	5.782	17.017	0	0	0	0.421	0	0.905	13.679	11.990	0.206	50	0.640	0.356	0.356	0.356
FH 049_13	0.4%	0.6	5.780	17.011	0	0	0	0.247	0	1.096	13.874	11.986	0.206	50	0.640	0.356	0.356	0.356
FH 050_13	0.4%	0.6	5.783	17.050	0.809	0	0	0.421	0	13.707	12.013	0.206	50	0.639	0.357	0.357	0.357	
FH 051_13	0.4%	0.6	5.783	17.049	0.984	0	0	0.248	0	13.707	12.013	0.206	50	0.639	0.358	0.358	0.358	
FH 052_13	0.4%	0.6	5.770	16.980	0.408	0.611	0	0.411	0	13.651	11.964	0.205	50	0.639	0.358	0.358	0.358	
FH 053_13	0.4%	0.6	5.725	17.196	0.486	0.728	0	0.245	0	13.545	11.871	0.204	50	0.639	0.358	0.358	0.358	
FH 058_13	0.4%	0.6	5.704	17.134	0.906	0.606	0	0.122	0	13.496	11.828	0.203	50	0.636	0.360	0.360	0.360	
FH 059_13	0.4%	0.6	5.753	17.280	0.000	1.100	0	0.123	0	13.611	11.929	0.205	50	0.637	0.360	0.360	0.360	
FH 060_13	0.4%	0.6	5.793	17.049	1.107	0.000	0	0.124	0	13.707	12.013	0.206	50	0.638	0.358	0.358	0.358	
FH 061_13	0.4%	0.6	5.722	17.186	0.818	0.548	0	0.122	0	13.537	11.864	0.204	50	0.638	0.359	0.359	0.359	
FH 064_13	0.8%	0.8	2.791	21.907	0.119	0	0	0	0	13.209	11.577	0.398	50	0.634	0.364	0.364	0.364	
FH 065_13	1.0%	0.8	2.786	21.863	0.118	0	0	0	0	13.183	11.554	0.496	50	0.638	0.360	0.360	0.360	
FH 066_13	1.2%	0.8	2.780	21.820	0.118	0	0	0	0	13.157	11.531	0.594	50	0.641	0.357	0.357	0.357	
FH 067_13	1.2%	0.9	1.354	23.908	0.115	0	0	0	0	12.814	11.230	0.579	50	0.633	0.386	0.386	0.386	
FH 068_13	1.2%	0.875	1.704	23.397	0.116	0	0	0	0	12.898	11.304	0.582	50	0.636	0.362	0.362	0.362	
FH 069_13	1.2%	0.85	2.058	22.878	0.117	0	0	0	0	12.983	11.378	0.586	50	0.640	0.358	0.358	0.358	
FH 070_13	1.2%	0.825	2.417	22.353	0.117	0	0	0	0	13.059	11.454	0.590	50	0.642	0.356	0.356	0.356	
FH 079_13	1.2%	0.925	1.009	24.413	0.114	0	0	0	0	12.731	11.158	0.575	50	0.628	0.370	0.370	0.370	
FH 080_13	1.2%	0.95	0.668	24.912	0.114	0	0	0	0	12.649	11.086	0.571	50	0.626	0.372	0.372	0.372	
FH 081_13	1.5%	0.8	2.772	21.756	0.118	0	0	0	0	13.118	11.497	0.740	50	0.645	0.353	0.353	0.353	
FH 082_13	2.0%	0.8	2.759	21.649	0.117	0	0	0	0	13.053	11.440	0.982	50	0.654	0.345	0.345	0.345	
TF 88/12	0.5%	0.7	8.632	39.518	0.244	0	0	0	0	27.227	23.866	0.512	100	0.635	0.364	0.364	0.364	
TF 93/12	0.6%	0.8	5.594	43.903	0.238	0	0	0	0	26.468	23.200	0.598	100	0.638	0.361	0.361	0.361	
TF 99/12	0.5%	0.7	8.632	39.518	0.244	0	0	0	0	27.227	23.866	0.512	100	0.632	0.366	0.366	0.366	

Fig. 39a/3

Experiment	y	Composition Ca _{1-x} Sr _x AlSiN ₃ yEu										Optical characterization		
		Sr/3N2	AlF3	SrF2	BaF2	Li2B4O7	LiBF4	Si3N4	AlN	Eu2O3	Total	CIE x	CIE y	
		Weights/g												
TF 100/12	0.6%	0.8	5.594	43.903	0.238	0	0	0	26.468	23.200	0.598	100	0.628	0.371
TF 103/12	0.5%	0.8	5.600	43.947	0.238	0	0	0	26.494	23.223	0.499	100	0.624	0.375
TF 131/12	0.6%	0.9	2.724	48.097	0.231	0	0	0	25.774	22.592	0.582	100	0.622	0.377
TF 132/12	0.7%	0.9	2.721	48.050	0.231	0	0	0	25.749	22.570	0.678	100	0.624	0.375
TF 133/12	0.8%	0.9	2.719	48.004	0.231	0	0	0	25.724	22.548	0.774	100	0.626	0.372
TF 161/12	1.0%	0.9	2.713	47.911	0.231	0	0	0	25.674	22.505	0.966	100	0.635	0.364
TF 162/12	1.2%	0.9	2.708	47.818	0.230	0	0	0	25.625	22.461	1.157	100	0.637	0.362
TF 163/12	0.6%	0.85	4.140	46.028	0.235	0	0	0	26.116	22.892	0.590	100	0.624	0.374
TF 168/12	0.9%	0.9	2.716	47.957	0.231	0	0	0	25.699	22.527	0.870	100	0.629	0.370
TF 169/12	1.2%	0.95	1.337	49.826	0.227	0	0	0	25.295	22.173	1.142	100	0.631	0.368

Fig. 39b/1

Fig. 39b/2

Fig. 39b/3

CONTINUATION OF TABLE 1 Ref. CA 02925738, filed 01/20											Continuation of Table 1																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																											
Element	Y	Z	Y₂	Z₂	Y₃	Z₃	Y₄	Z₄	Y₅	Z₅	Y₆	Z₆	Y₇	Z₇	Y₈	Z₈	Y₉	Z₉	Y₁₀	Z₁₀	Y₁₁	Z₁₁	Y₁₂	Z₁₂	Y₁₃	Z₁₃	Y₁₄	Z₁₄	Y₁₅	Z₁₅	Y₁₆	Z₁₆	Y₁₇	Z₁₇	Y₁₈	Z₁₈	Y₁₉	Z₁₉	Y₂₀	Z₂₀	Y₂₁	Z₂₁	Y₂₂	Z₂₂	Y₂₃	Z₂₃	Y₂₄	Z₂₄	Y₂₅	Z₂₅	Y₂₆	Z₂₆	Y₂₇	Z₂₇	Y₂₈	Z₂₈	Y₂₉	Z₂₉	Y₃₀	Z₃₀	Y₃₁	Z₃₁	Y₃₂	Z₃₂	Y₃₃	Z₃₃	Y₃₄	Z₃₄	Y₃₅	Z₃₅	Y₃₆	Z₃₆	Y₃₇	Z₃₇	Y₃₈	Z₃₈	Y₃₉	Z₃₉	Y₄₀	Z₄₀	Y₄₁	Z₄₁	Y₄₂	Z₄₂	Y₄₃	Z₄₃	Y₄₄	Z₄₄	Y₄₅	Z₄₅	Y₄₆	Z₄₆	Y₄₇	Z₄₇	Y₄₈	Z₄₈	Y₄₉	Z₄₉	Y₅₀	Z₅₀	Y₅₁	Z₅₁	Y₅₂	Z₅₂	Y₅₃	Z₅₃	Y₅₄	Z₅₄	Y₅₅	Z₅₅	Y₅₆	Z₅₆	Y₅₇	Z₅₇	Y₅₈	Z₅₈	Y₅₉	Z₅₉	Y₆₀	Z₆₀	Y₆₁	Z₆₁	Y₆₂	Z₆₂	Y₆₃	Z₆₃	Y₆₄	Z₆₄	Y₆₅	Z₆₅	Y₆₆	Z₆₆	Y₆₇	Z₆₇	Y₆₈	Z₆₈	Y₆₉	Z₆₉	Y₇₀	Z₇₀	Y₇₁	Z₇₁	Y₇₂	Z₇₂	Y₇₃	Z₇₃	Y₇₄	Z₇₄	Y₇₅	Z₇₅	Y₇₆	Z₇₆	Y₇₇	Z₇₇	Y₇₈	Z₇₈	Y₇₉	Z₇₉	Y₈₀	Z₈₀	Y₈₁	Z₈₁	Y₈₂	Z₈₂	Y₈₃	Z₈₃	Y₈₄	Z₈₄	Y₈₅	Z₈₅	Y₈₆	Z₈₆	Y₈₇	Z₈₇	Y₈₈	Z₈₈	Y₈₉	Z₈₉	Y₉₀	Z₉₀	Y₉₁	Z₉₁	Y₉₂	Z₉₂	Y₉₃	Z₉₃	Y₉₄	Z₉₄	Y₉₅	Z₉₅	Y₉₆	Z₉₆	Y₉₇	Z₉₇	Y₉₈	Z₉₈	Y₉₉	Z₉₉	Y₁₀₀	Z₁₀₀	Y₁₀₁	Z₁₀₁	Y₁₀₂	Z₁₀₂	Y₁₀₃	Z₁₀₃	Y₁₀₄	Z₁₀₄	Y₁₀₅	Z₁₀₅	Y₁₀₆	Z₁₀₆	Y₁₀₇	Z₁₀₇	Y₁₀₈	Z₁₀₈	Y₁₀₉	Z₁₀₉	Y₁₁₀	Z₁₁₀	Y₁₁₁	Z₁₁₁	Y₁₁₂	Z₁₁₂	Y₁₁₃	Z₁₁₃	Y₁₁₄	Z₁₁₄	Y₁₁₅	Z₁₁₅	Y₁₁₆	Z₁₁₆	Y₁₁₇	Z₁₁₇	Y₁₁₈	Z₁₁₈	Y₁₁₉	Z₁₁₉	Y₁₂₀	Z₁₂₀	Y₁₂₁	Z₁₂₁	Y₁₂₂	Z₁₂₂	Y₁₂₃	Z₁₂₃	Y₁₂₄	Z₁₂₄	Y₁₂₅	Z₁₂₅	Y₁₂₆	Z₁₂₆	Y₁₂₇	Z₁₂₇	Y₁₂₈	Z₁₂₈	Y₁₂₉	Z₁₂₉	Y₁₃₀	Z₁₃₀	Y₁₃₁	Z₁₃₁	Y₁₃₂	Z₁₃₂	Y₁₃₃	Z₁₃₃	Y₁₃₄	Z₁₃₄	Y₁₃₅	Z₁₃₅	Y₁₃₆	Z₁₃₆	Y₁₃₇	Z₁₃₇	Y₁₃₈	Z₁₃₈	Y₁₃₉	Z₁₃₉	Y₁₄₀	Z₁₄₀	Y₁₄₁	Z₁₄₁	Y₁₄₂	Z₁₄₂	Y₁₄₃	Z₁₄₃	Y₁₄₄	Z₁₄₄	Y₁₄₅	Z₁₄₅	Y₁₄₆	Z₁₄₆	Y₁₄₇	Z₁₄₇	Y₁₄₈	Z₁₄₈	Y₁₄₉	Z₁₄₉	Y₁₅₀	Z₁₅₀	Y₁₅₁	Z₁₅₁	Y₁₅₂	Z₁₅₂	Y₁₅₃	Z₁₅₃	Y₁₅₄	Z₁₅₄	Y₁₅₅	Z₁₅₅	Y₁₅₆	Z₁₅₆	Y₁₅₇	Z₁₅₇	Y₁₅₈	Z₁₅₈	Y₁₅₉	Z₁₅₉	Y₁₆₀	Z₁₆₀	Y₁₆₁	Z₁₆₁	Y₁₆₂	Z₁₆₂	Y₁₆₃	Z₁₆₃	Y₁₆₄	Z₁₆₄	Y₁₆₅	Z₁₆₅	Y₁₆₆	Z₁₆₆	Y₁₆₇	Z₁₆₇	Y₁₆₈	Z₁₆₈	Y₁₆₉	Z₁₆₉	Y₁₇₀	Z₁₇₀	Y₁₇₁	Z₁₇₁	Y₁₇₂	Z₁₇₂	Y₁₇₃	Z₁₇₃	Y₁₇₄	Z₁₇₄	Y₁₇₅	Z₁₇₅	Y₁₇₆	Z₁₇₆	Y₁₇₇	Z₁₇₇	Y₁₇₈	Z₁₇₈	Y₁₇₉	Z₁₇₉	Y₁₈₀	Z₁₈₀	Y₁₈₁	Z₁₈₁	Y₁₈₂	Z₁₈₂	Y₁₈₃	Z₁₈₃	Y₁₈₄	Z₁₈₄	Y₁₈₅	Z₁₈₅	Y₁₈₆	Z₁₈₆	Y₁₈₇	Z₁₈₇	Y₁₈₈	Z₁₈₈	Y₁₈₉	Z₁₈₉	Y₁₉₀	Z₁₉₀	Y₁₉₁	Z₁₉₁	Y₁₉₂	Z₁₉₂	Y₁₉₃	Z₁₉₃	Y₁₉₄	Z₁₉₄	Y₁₉₅	Z₁₉₅	Y₁₉₆	Z₁₉₆	Y₁₉₇	Z₁₉₇	Y₁₉₈	Z₁₉₈	Y₁₉₉	Z₁₉₉	Y₂₀₀	Z₂₀₀	Y₂₀₁	Z₂₀₁	Y₂₀₂	Z₂₀₂	Y₂₀₃	Z₂₀₃	Y₂₀₄	Z₂₀₄	Y₂₀₅	Z₂₀₅	Y₂₀₆	Z₂₀₆	Y₂₀₇	Z₂₀₇	Y₂₀₈	Z₂₀₈	Y₂₀₉	Z₂₀₉	Y₂₁₀	Z₂₁₀	Y₂₁₁	Z₂₁₁	Y₂₁₂	Z₂₁₂	Y₂₁₃	Z₂₁₃	Y₂₁₄	Z₂₁₄	Y₂₁₅	Z₂₁₅	Y₂₁₆	Z₂₁₆	Y₂₁₇	Z₂₁₇	Y₂₁₈	Z₂₁₈	Y₂₁₉	Z₂₁₉	Y₂₂₀	Z₂₂₀	Y₂₂₁	Z₂₂₁	Y₂₂₂	Z₂₂₂	Y₂₂₃	Z₂₂₃	Y₂₂₄	Z₂₂₄	Y₂₂₅	Z₂₂₅	Y₂₂₆	Z₂₂₆	Y₂₂₇	Z₂₂₇	Y₂₂₈	Z₂₂₈	Y₂₂₉	Z₂₂₉	Y₂₃₀	Z₂₃₀	Y₂₃₁	Z₂₃₁	Y₂₃₂	Z₂₃₂	Y₂₃₃	Z₂₃₃	Y₂₃₄	Z₂₃₄	Y₂₃₅	Z₂₃₅	Y₂₃₆	Z₂₃₆	Y₂₃₇	Z₂₃₇	Y₂₃₈	Z₂₃₈	Y₂₃₉	Z₂₃₉	Y₂₄₀	Z₂₄₀	Y₂₄₁	Z₂₄₁	Y₂₄₂	Z₂₄₂	Y₂₄₃	Z₂₄₃	Y₂₄₄	Z₂₄₄	Y₂₄₅	Z₂₄₅	Y₂₄₆	Z₂₄₆	Y₂₄₇	Z₂₄₇	Y₂₄₈	Z₂₄₈	Y₂₄₉	Z₂₄₉	Y₂₅₀	Z₂₅₀	Y₂₅₁	Z₂₅₁	Y₂₅₂	Z₂₅₂	Y₂₅₃	Z₂₅₃	Y₂₅₄	Z₂₅₄	Y₂₅₅	Z₂₅₅	Y₂₅₆	Z₂₅₆	Y₂₅₇	Z₂₅₇	Y₂₅₈	Z₂₅₈	Y₂₅₉	Z₂₅₉	Y₂₆₀	Z₂₆₀	Y₂₆₁	Z₂₆₁	Y₂₆₂	Z₂₆₂	Y₂₆₃	Z₂₆₃	Y₂₆₄	Z₂₆₄	Y₂₆₅	Z₂₆₅	Y₂₆₆	Z₂₆₆	Y₂₆₇	Z₂₆₇	Y₂₆₈	Z₂₆₈	Y₂₆₉	Z₂₆₉	Y₂₇₀	Z₂₇₀	Y₂₇₁	Z₂₇₁	Y₂₇₂	Z₂₇₂	Y₂₇₃	Z₂₇₃	Y₂₇₄	Z₂₇₄	Y₂₇₅	Z₂₇₅	Y₂₇₆	Z₂₇₆	Y₂₇₇	Z₂₇₇	Y₂₇₈	Z₂₇₈	Y₂₇₉	Z₂₇₉	Y₂₈₀	Z₂₈₀	Y₂₈₁	Z₂₈₁	Y₂₈₂	Z₂₈₂	Y₂₈₃	Z₂₈₃	Y₂₈₄	Z₂₈₄	Y₂₈₅	Z₂₈₅	Y₂₈₆	Z₂₈₆	Y₂₈₇	Z₂₈₇	Y₂₈₈	Z₂₈₈	Y₂₈₉	Z₂₈₉	Y₂₉₀	Z₂₉₀	Y₂₉₁	Z₂₉₁	Y₂₉₂	Z₂₉₂	Y₂₉₃	Z₂₉₃	Y₂₉₄	Z₂₉₄	Y₂₉₅	Z₂₉₅	Y₂₉₆	Z₂₉₆	Y₂₉₇	Z₂₉₇	Y₂₉₈	Z₂₉₈	Y₂₉₉	Z₂₉₉	Y₃₀₀	Z₃₀₀	Y₃₀₁	Z₃₀₁	Y₃₀₂	Z₃₀₂	Y₃₀₃	Z₃₀₃	Y₃₀₄	Z₃₀₄	Y₃₀₅	Z₃₀₅	Y₃₀₆	Z₃₀₆	Y₃₀₇	Z₃₀₇	Y₃₀₈	Z₃₀₈	Y₃₀₉	Z₃₀₉	Y₃₁₀	Z₃₁₀	Y₃₁₁	Z₃₁₁	Y₃₁₂	Z₃₁₂	Y₃₁₃	Z₃₁₃	Y₃₁₄	Z₃₁₄	Y₃₁₅	Z₃₁₅	Y₃₁₆	Z₃₁₆	Y₃₁₇	Z₃₁₇	Y₃₁₈	Z₃₁₈	Y₃₁₉	Z₃₁₉	Y₃₂₀	Z₃₂₀	Y₃₂₁	Z₃₂₁	Y₃₂₂	Z₃₂₂	Y₃₂₃	Z₃₂₃	Y₃₂₄	Z₃₂₄	Y₃₂₅	Z₃₂₅	Y₃₂₆	Z₃₂₆	Y₃₂₇	Z₃₂₇	Y₃₂₈	Z₃₂₈	Y₃₂₉	Z₃₂₉	Y₃₃₀	Z₃₃₀	Y₃₃₁	Z₃₃₁	Y₃₃₂	Z₃₃₂	Y₃₃₃	Z₃₃₃	Y₃₃₄	Z₃₃₄	Y₃₃₅	Z₃₃₅	Y₃₃₆	Z₃₃₆	Y₃₃₇	Z₃₃₇	Y₃₃₈	Z₃₃₈	Y₃₃₉	Z₃₃₉	Y₃₄₀	Z₃₄₀	Y₃₄₁	Z₃₄₁	Y₃₄₂	Z₃₄₂	Y₃₄₃	Z₃₄₃	Y₃₄₄	Z₃₄₄	Y₃₄₅	Z₃₄₅	Y₃₄₆	Z₃₄₆	Y₃₄₇	Z₃₄₇	Y₃₄₈	Z₃₄₈	Y₃₄₉	Z₃₄₉	Y₃₅₀	Z₃₅₀	Y₃₅₁	Z₃₅₁	Y₃₅₂	Z₃₅₂	Y₃₅₃	Z₃₅₃	Y₃₅₄	Z₃₅₄	Y₃₅₅	Z₃₅₅	Y₃₅₆	Z₃₅₆	Y₃₅₇	Z₃₅₇	Y₃₅₈	Z₃₅₈	Y₃₅₉	Z₃₅₉	Y₃₆₀	Z₃₆₀	Y₃₆₁	Z₃₆₁	Y₃₆₂	Z₃₆₂	Y₃₆₃	Z₃₆₃	Y₃₆₄	Z₃₆₄	Y₃₆₅	Z₃₆₅	Y₃₆₆	Z₃₆₆	Y₃₆₇	Z₃₆₇	Y₃₆₈	Z₃₆₈	Y₃₆₉	Z₃₆₉	Y₃₇₀	Z₃₇₀	Y₃₇₁	Z₃₇₁	Y₃₇₂	Z₃₇₂	Y₃₇₃	Z₃₇₃	Y₃₇₄	Z₃₇₄	Y₃₇₅	Z₃₇₅	Y₃₇₆	Z₃₇₆	Y₃₇₇	Z₃₇₇	Y₃₇₈	Z₃₇₈	Y₃₇₉	Z₃₇₉	Y₃₈₀	Z₃₈₀	Y₃₈₁	Z₃₈₁	Y₃₈₂	Z₃₈₂	Y₃₈₃	Z₃₈₃	Y₃₈₄	Z₃₈₄	Y₃₈₅	Z₃₈₅	Y₃₈₆	Z₃₈₆	Y₃₈₇	Z₃₈₇	Y₃₈₈	Z₃₈₈	Y₃₈₉	Z₃₈₉	Y₃₉₀	Z₃₉₀	Y₃₉₁	Z₃₉₁	Y₃₉₂	Z₃₉₂	Y₃₉₃	Z₃₉₃	Y₃₉₄	Z₃₉₄	Y₃₉₅	Z₃₉₅	Y₃₉₆	Z₃₉₆	Y₃₉₇	Z₃₉₇	Y₃₉₈	Z₃₉₈	Y₃₉₉	Z₃₉₉	Y₄₀₀	Z₄₀₀	Y₄₀₁	Z₄₀₁	Y₄₀₂	Z₄₀₂	Y₄₀₃	Z₄₀₃	Y₄₀₄	Z₄₀₄	Y₄₀₅	Z₄₀₅	Y₄₀₆	Z₄₀₆	Y₄₀₇	Z₄₀₇	Y₄₀₈	Z₄₀₈	Y₄₀₉	Z₄₀₉	Y₄₁₀	Z₄₁₀	Y₄₁₁	Z₄₁₁
<th

Fig. 39b/4

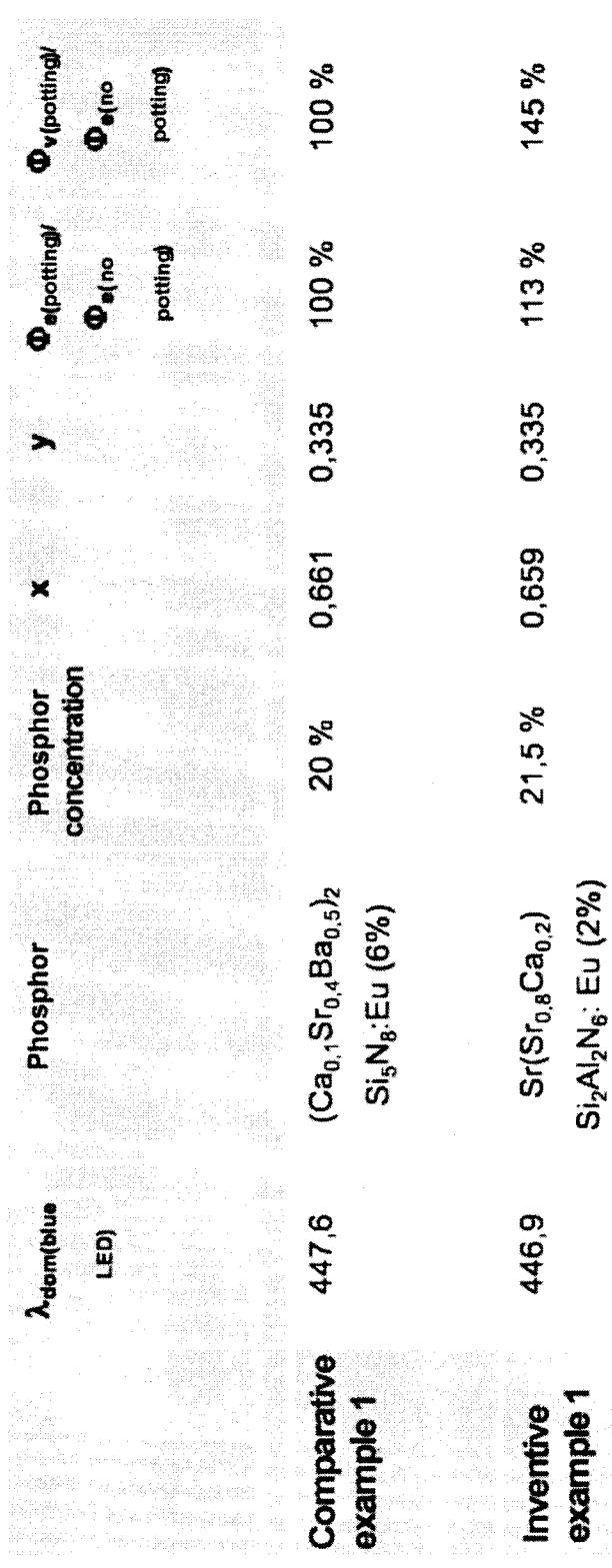
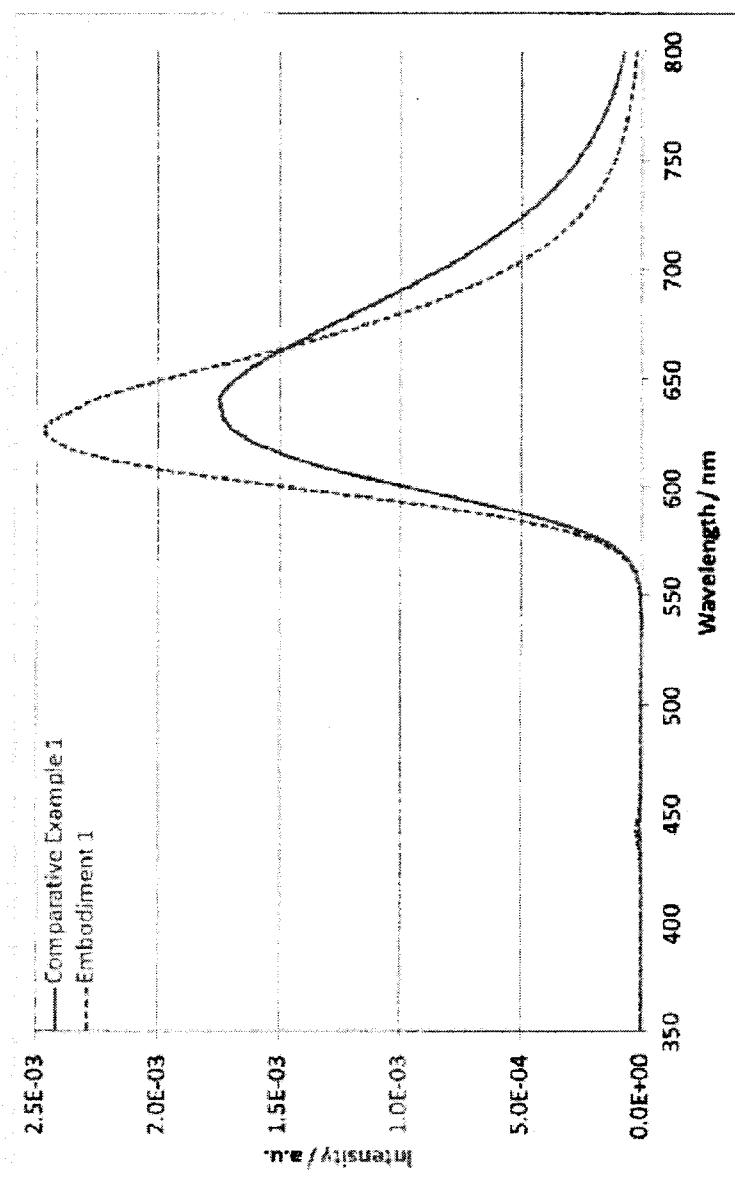


Fig. 40a

Fig. 40b



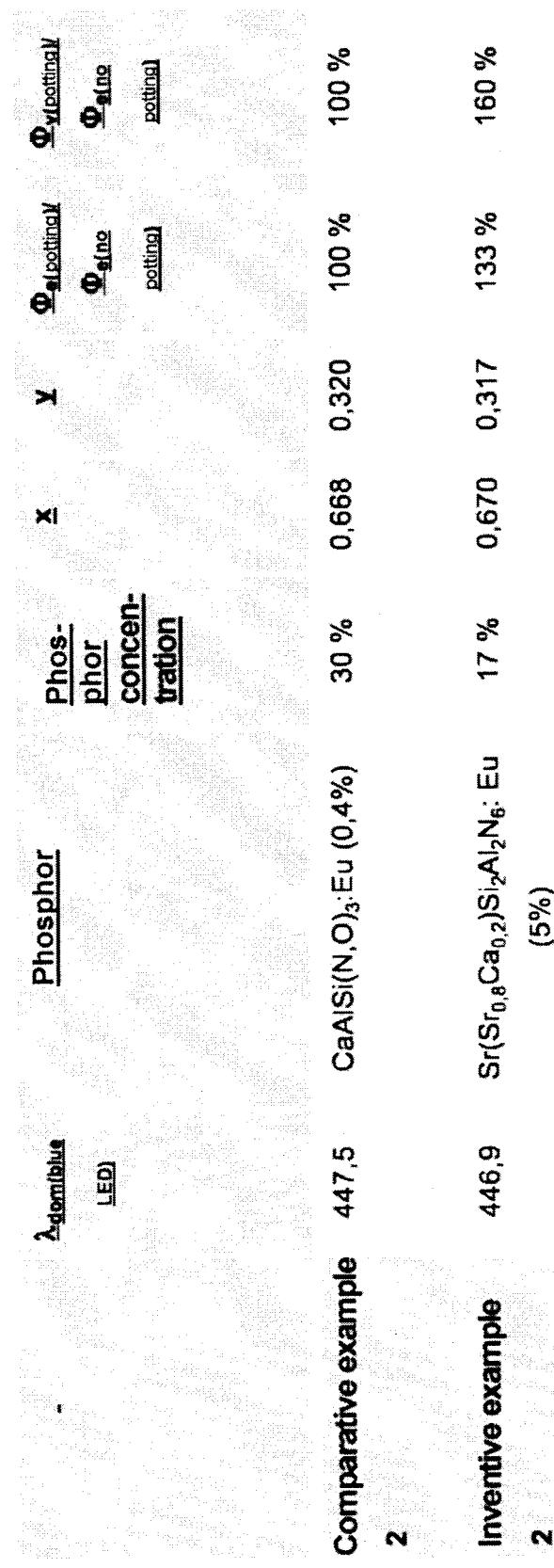


Fig. 41a

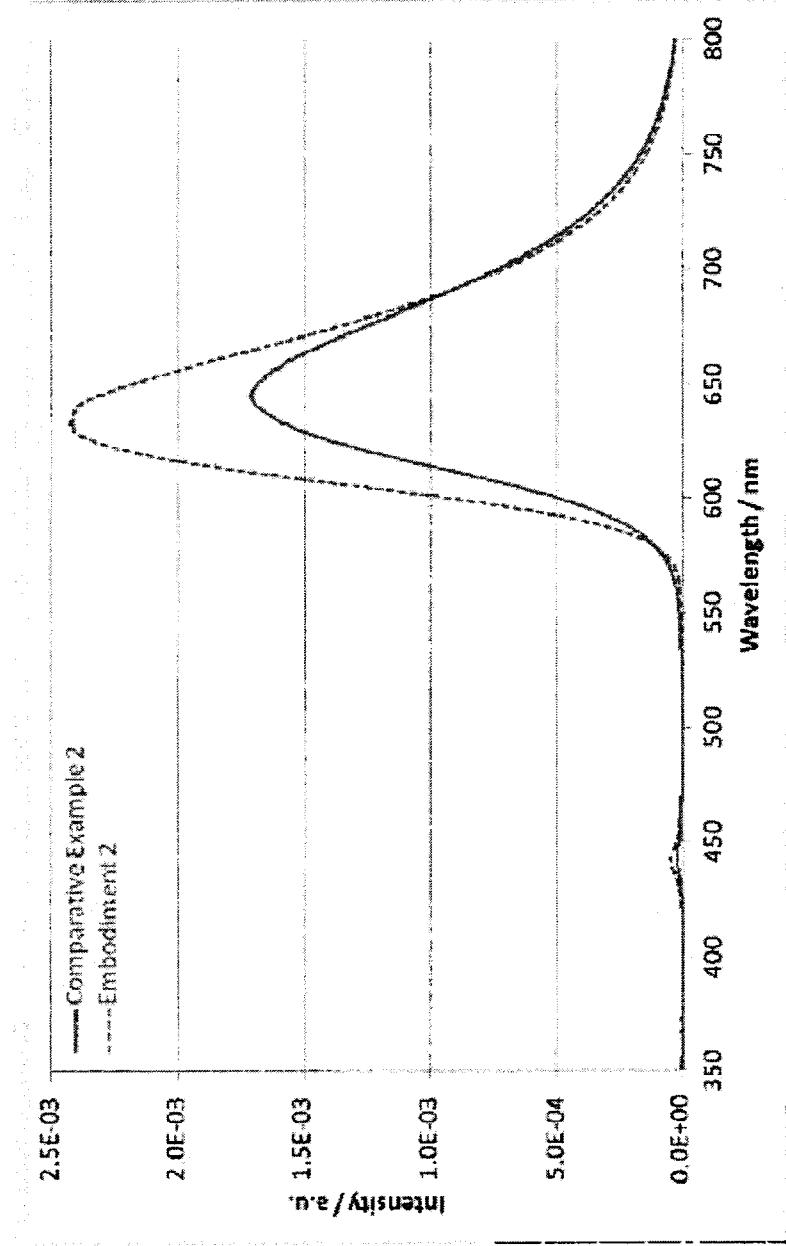


Fig. 41b

Fig. 42a

First phosphor $\lambda_{\text{dom(blue)}}$ (LED)	Second phosphor concentration	Ratio of phosphor concentration	Phosphor concentration		
			$\Phi_{\text{e}(pott)}$	$\Phi_{\text{e}(no pott)}$	$\Phi_{\text{e}(no pott)}$
Comparative example 3	447,6 ($\text{Ca}_{0,1}\text{Sr}_{0,4}\text{Ba}_{0,5}$) ₂ $\text{Si}_3\text{N}_4\text{:Eu}$ (6%)	20 %	0,661	0,335	100 %
Inventive example 3	447,0 CaAlSi(N,O) ₃ :Eu (0,4%)	0,5/1 $\text{Sr}(\text{Sr}_{0,8}\text{Ca}_{0,2})\text{Si}_2$ $\text{Al}_2\text{N}_6\text{:Eu}$ (2%)	21,5 %	0,661	0,330
					104 %
					118 %

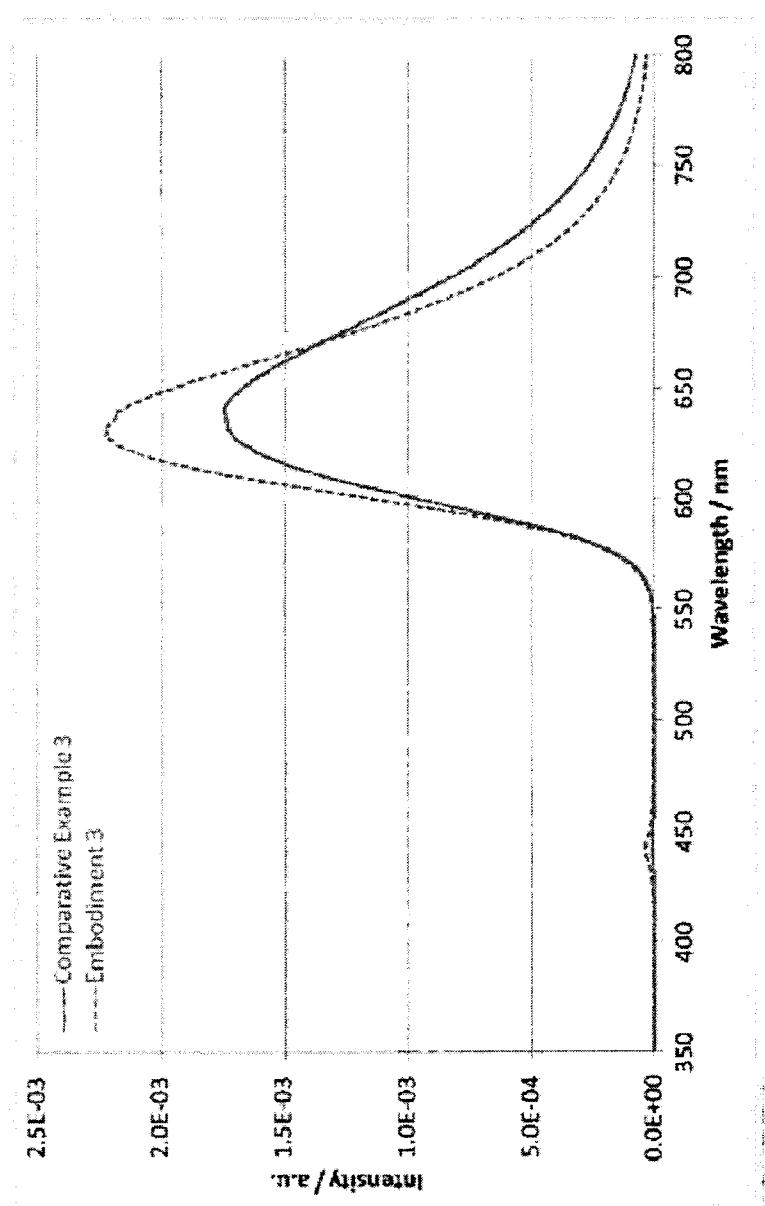


Fig. 42b

Fig. 43a

	$\lambda_{\text{demi/blue}}$ (LED)	First phosphor	Second phosphor	Ratio of phosphors	Phosphor concentration	$\Phi_{\text{e}(no\text{-}potting)}$	$\Phi_{\text{e}(potting)}$
Comparative example 4	447,1	CaAlSi(N ₃ O) ₃ :Eu (0,4%)	(Ca _{0,1} Si _{0,4} Ba _{0,5}) ₂ Si ₅ N ₈ :Eu (4%)	1,7/1	21,5 %	0,661	0,329
Inventive example 4	447,0	CaAlSi(N ₃ O) ₃ :Eu (0,4%)	Sr(Sr _{0,8} Ca _{0,2})Si ₂ Al ₂ N ₆ :Eu (2%)	0,5/1	21,5 %	0,661	0,330

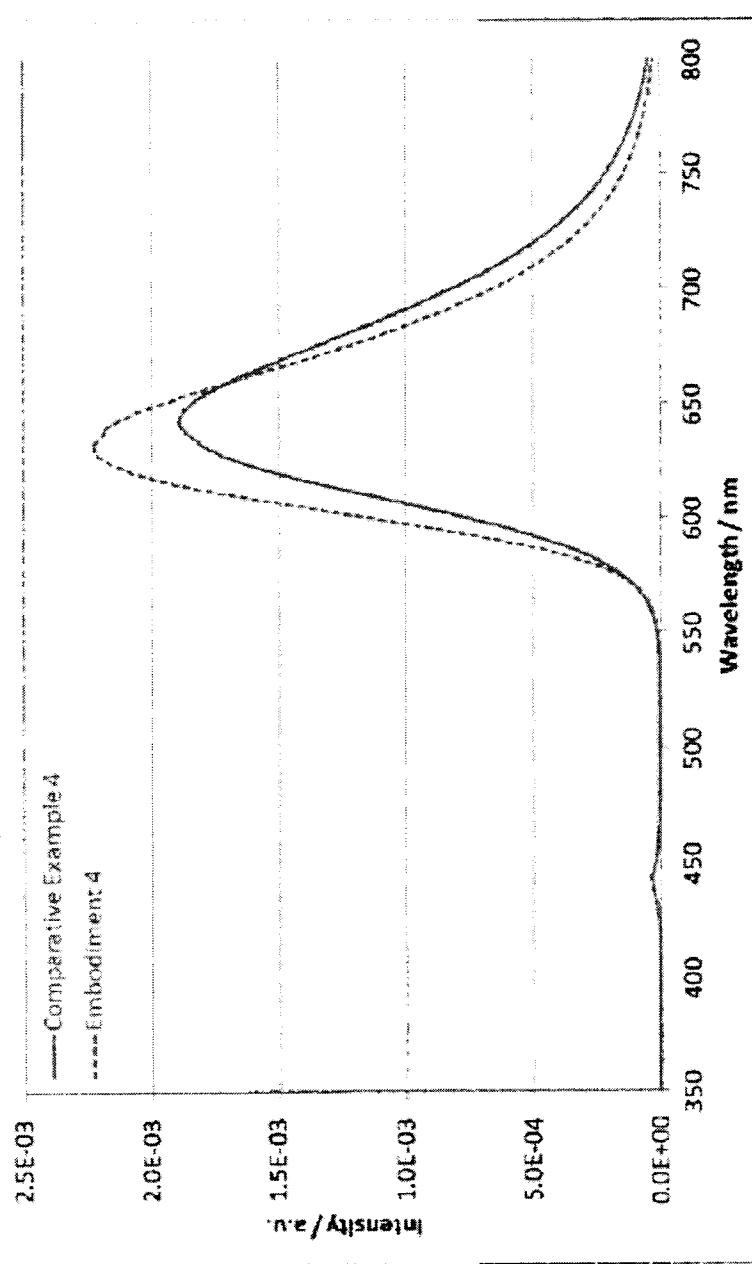


Fig. 43b

Fig. 44a

	$\lambda_{\text{dom}}(\text{nm})$	First phosphor LED)	Second phosphor phosphors	Ratio of phos- phor phos- phors	Phos- phor concen- tration	$\Phi_{\text{out}}(\text{pot-}ting)$	$\Phi_{\text{out}}(\text{no}potting)$
Comparative example 5	444,7	$(\text{Y}_{0,957}\text{Ce}_{0,043})_3\text{Al}_5\text{O}_{12}$	$(\text{Ca}_{0,1}\text{Sr}_{0,4}\text{Ba}_{0,5})_2$ $\text{Si}_3\text{N}_4\text{:Eu (4%)}$	5,9/1	41 %	0,561	0,430
Inventive example 5	444,4	$(\text{Y}_{0,957}\text{Ce}_{0,043})_3\text{Al}_5\text{O}_{12}$	$\text{Sr}(\text{Sr}_{0,6}\text{Ca}_{0,2})\text{Si}_2$ $\text{Al}_2\text{N}_6\text{:Eu (2%)}$	12,1/1	41 %	0,561	0,430
Inventive example 6	444,4	$(\text{Y}_{0,957}\text{Ce}_{0,043})_3\text{Al}_5\text{O}_{12}$	$\text{Sr}(\text{Sr}_{0,88}\text{Ca}_{0,14})\text{Si}$ $\text{Al}_2\text{N}_6\text{:Eu (0,8%)}$	5/1	39 %	0,561	0,428

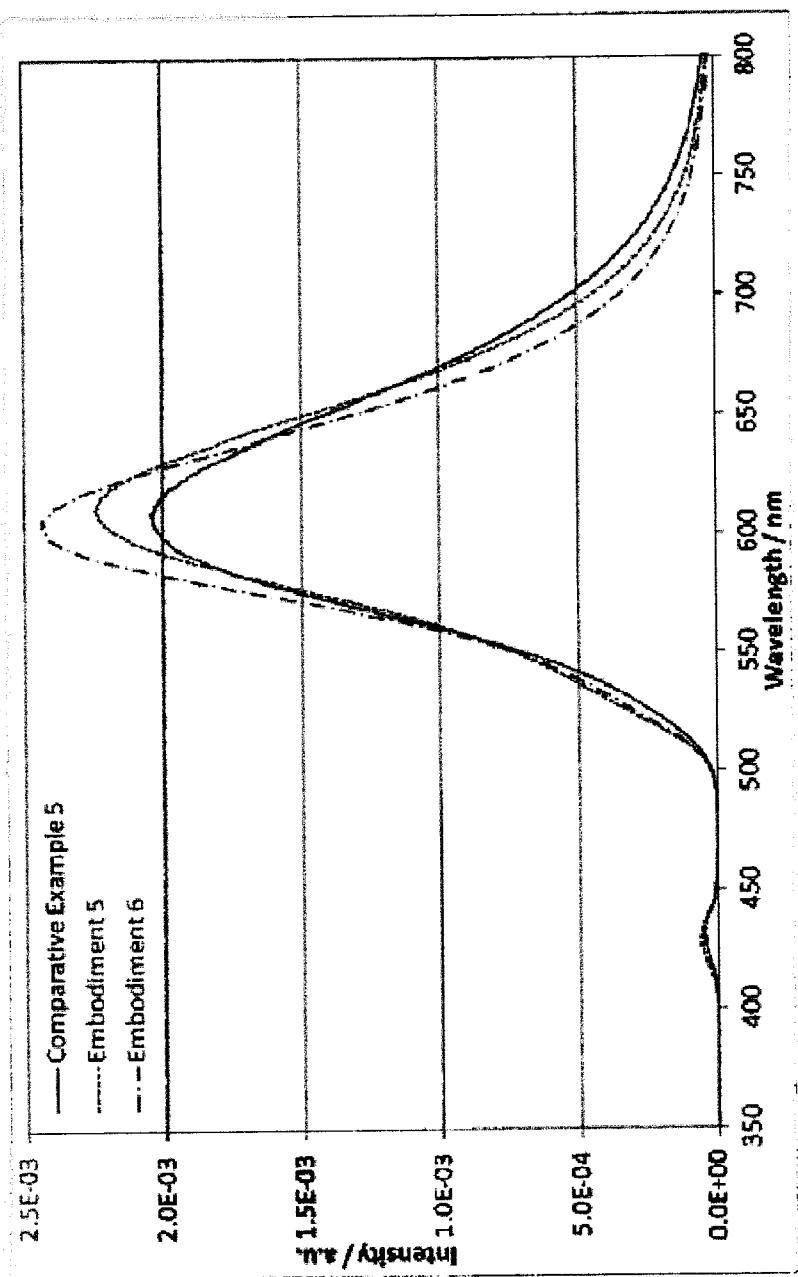
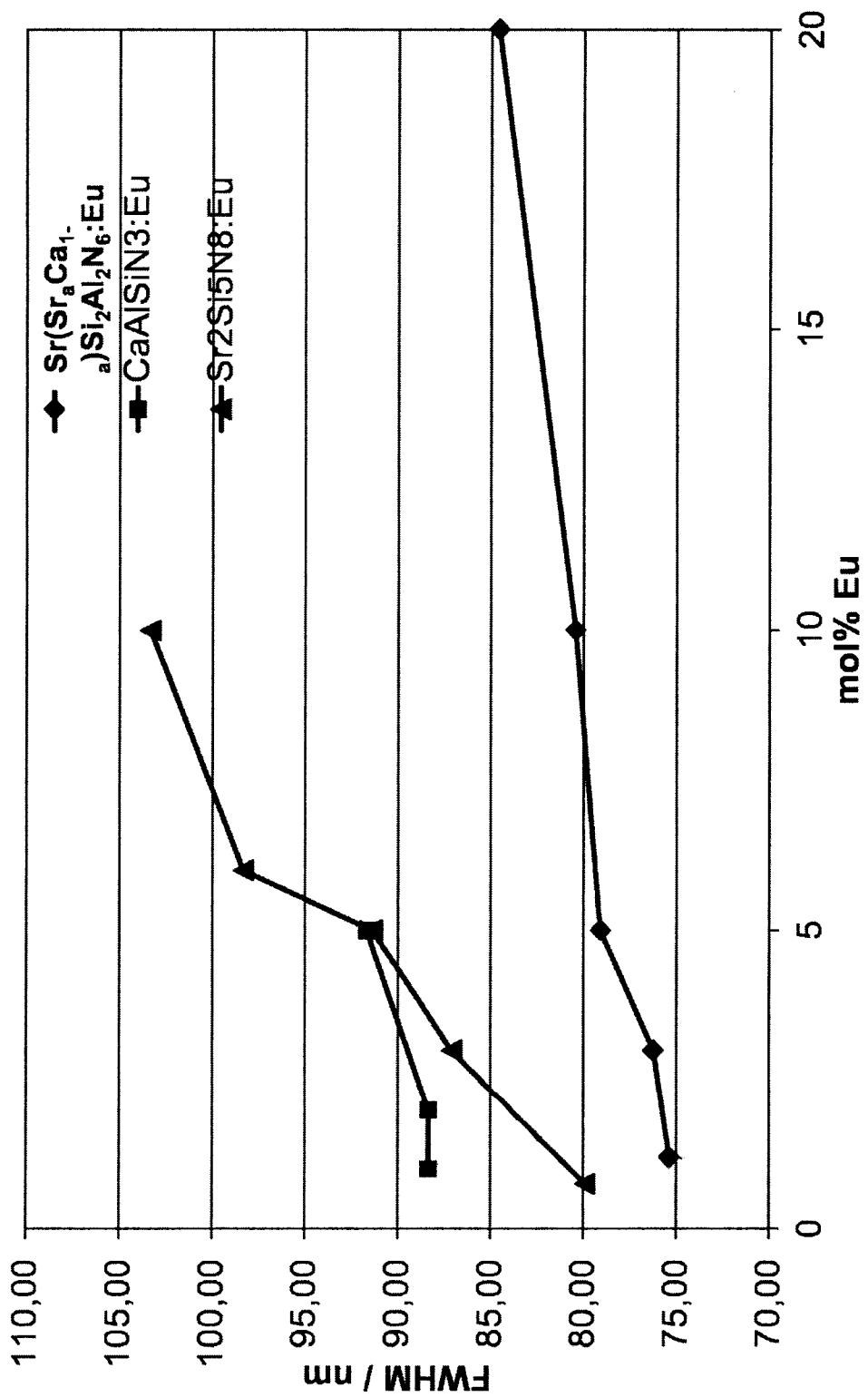


Fig. 44b

Fig. 45a

Example	λ_{dom} (blue LED)	Composition of red phosphor (% by wt.)	Phosphor concentration (% by wt.)	λ_{dom}	CIE-x	CIE-y
Inventive ex. 1	446,9	$\text{Sr}(\text{Sr}_{0,86}\text{Ca}_{0,14})\text{Si}_2$ Al_2N_6 ; Eu (10% Eu)	15,5%	621,2 nm	0,682	0,307
Inventive ex. 2	446,9	$\text{Sr}(\text{Sr}_{0,86}\text{Ca}_{0,14})\text{Si}_2$ Al_2N_6 ; Eu (10% Eu)	17,0%	621,2 nm	0,690	0,306
Inventive ex. 3	446,9	$\text{Sr}(\text{Sr}_{0,86}\text{Ca}_{0,14})\text{Si}_2$ Al_2N_6 ; Eu (20% Eu)	19,3%	632,3 nm	0,698	0,291

Fig. 45b



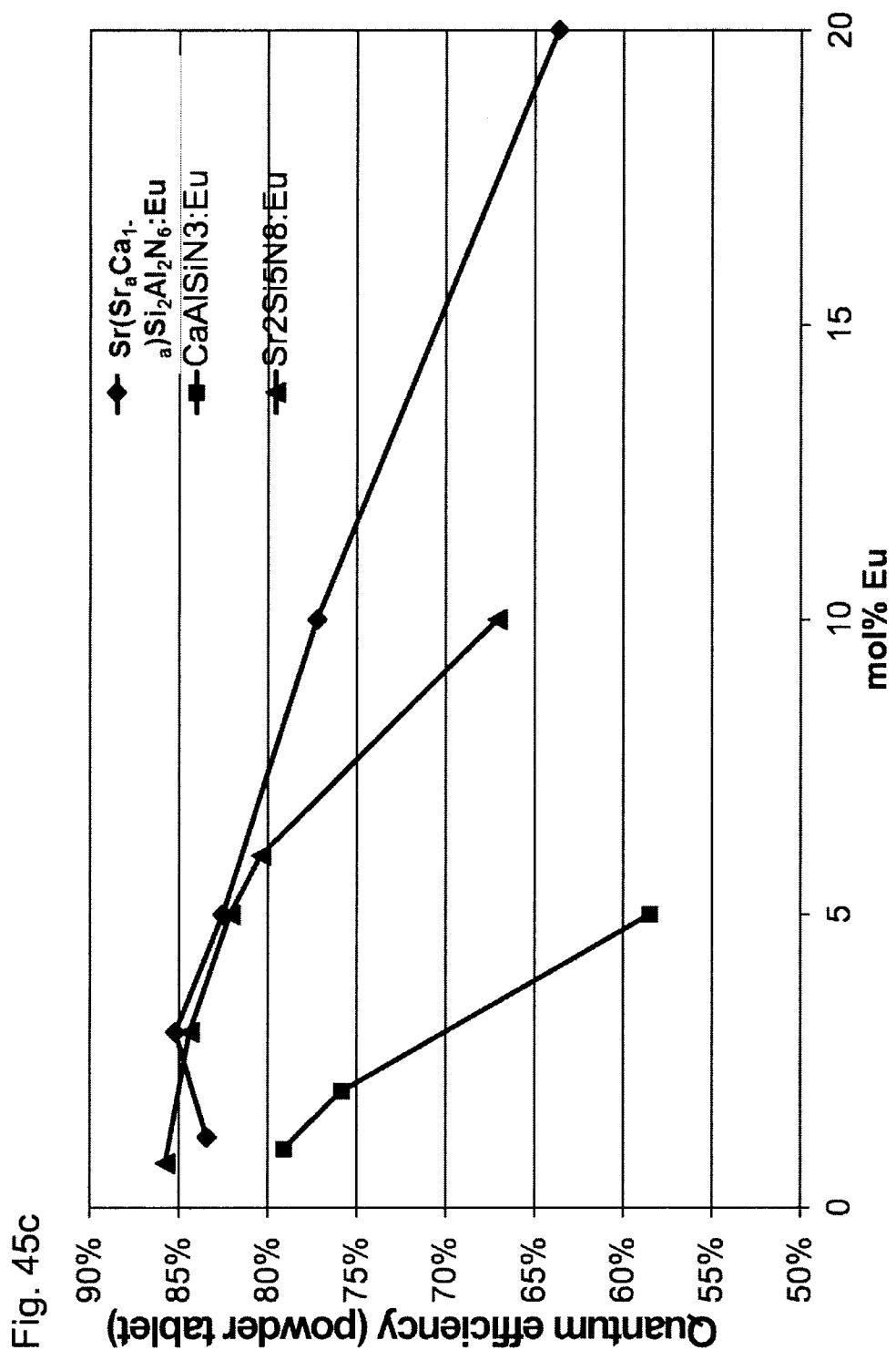


Fig. 45d

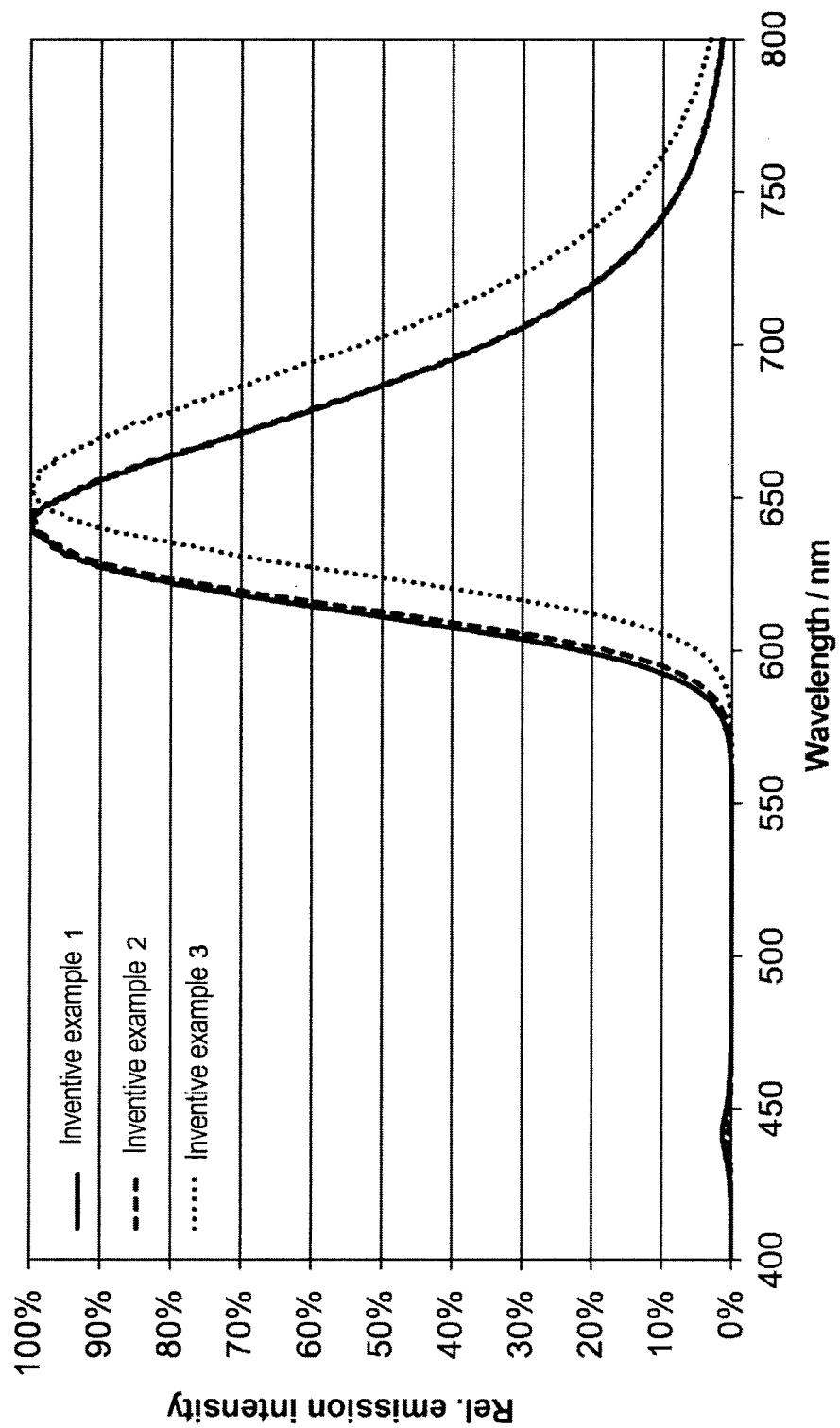


Fig. 46a

Example	$\lambda_{\text{emit}}(\text{blue LED})$ at 350 mA	Composition of yellow phosphor	Composition of red phosphor	Total phosphor conc. (% by wt.)	Yellow/red ratio
Comp. ex. 1	444.5	(Lu _{0.8} Y _{0.2}) ₃ Al ₅ O ₁₂ ; Ce (4.5% Ce)	(Ca _{0.025} Sr _{0.475} Ba _{0.5}) ₂ Si ₃ N ₈ ;Eu (2.5% Eu)	19.9%	7.0
Comp. ex. 2	444.6	(Lu _{0.85} Y _{0.15}) ₃ Al ₅ O ₁₂ ; Ce (3% Ce)	(Sr _{0.5} Ba _{0.5}) ₂ Si ₃ N ₈ ;Eu (1% Eu)	17.7%	3.0
Inventive ex. 1	444.6	(Lu _{0.85} Y _{0.15}) ₃ Al ₅ O ₁₂ ; Ce (3% Ce)	Sr(Sr _{0.86} Ca _{0.14})Si ₂ Al ₂ N ₆ ; Eu (0.4% Eu)	17.7%	3.0
Inventive ex. 2	444.4	(Lu _{0.85} Y _{0.15}) ₃ Al ₅ O ₁₂ ; Ce (3% Ce)	Sr(Sr _{0.86} Ca _{0.14})Si ₂ Al ₂ N ₆ ; Eu (1.2% Eu)	17.7%	8.1
Inventive ex. 3	444.5	Y ₃ (Al _{0.70} Ga _{0.30}) ₅ O ₁₂ ; Ce (3% Ce)	Sr(Sr _{0.86} Ca _{0.14})Si ₂ Al ₂ N ₆ ; Eu (1.2% Eu)	9.8%	4.0
Inventive ex. 4	444.5	Y _{0.3} (Al _{0.75} Ga _{0.25}) ₅ O ₁₂ ; Ce (3% Ce)	Sr(Sr _{0.86} Ca _{0.14})Si ₂ Al ₂ N ₆ ; Eu (1.2% Eu)	9.8%	4.2

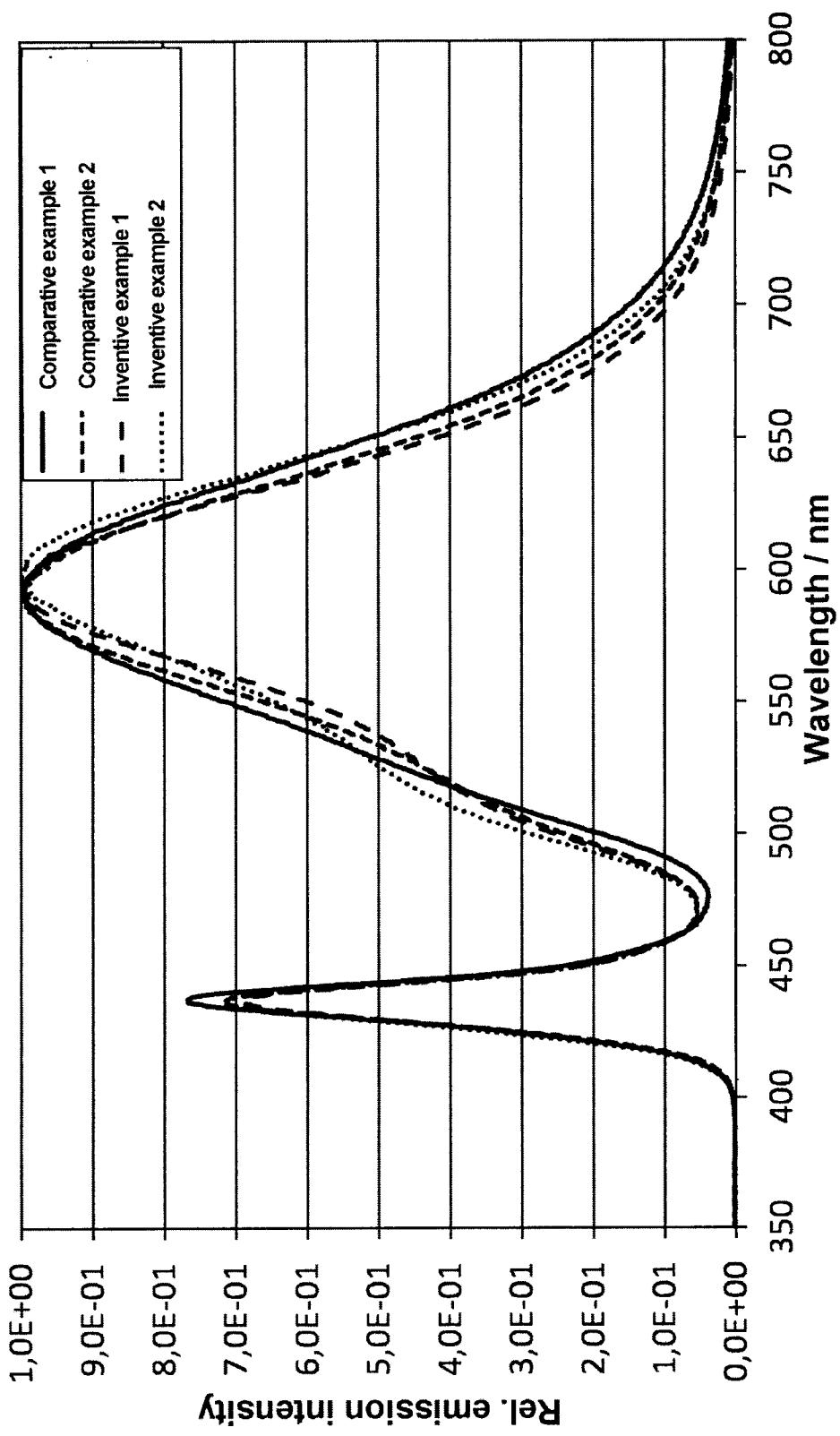
Example	CIE-x	CIE-y	CRI	Rel. Φ filled	
				potting/ Φ clear potting)	potting/ Φ clear potting) (color point corr.)
Comp. ex. 1	0,435	0,403	69	100%	100%
Comp. ex. 2	0,433	0,405	69	104,1%	103,4%
Inventive ex. 1	0,435	0,401	70	103%	103,4%
Inventive ex. 2	0,434	0,402	76	100,6%	100,7%
Inventive ex. 3	0,435	0,402	75	102,1%	102,4%
Inventive ex. 4	0,434	0,403	74	101,2%	101,1%

Fig. 46b

Fig. 46c

Example	CIE-x	CIE-y	CRI	Rel. $\Phi_{v\text{filled}}$ potting/ Φ_{clear} potting)	Rel. $\Phi_{v\text{filled}}$ potting/ Φ_{clear} potting)
Comp. ex. 1	0,430	0,401	69	100%	100%
Comp. ex. 2	0,429	0,402	69	104,4%	104,6%
Inventive ex. 1	0,433	0,399	71	104,1%	105,4%
Inventive ex. 2	0,430	0,400	75	101,6%	101,9%
Inventive ex. 3	0,432	0,400	75	104,4%	105,4%
Inventive ex. 4	0,431	0,401	74	102,7%	103,5%

Fig. 46d



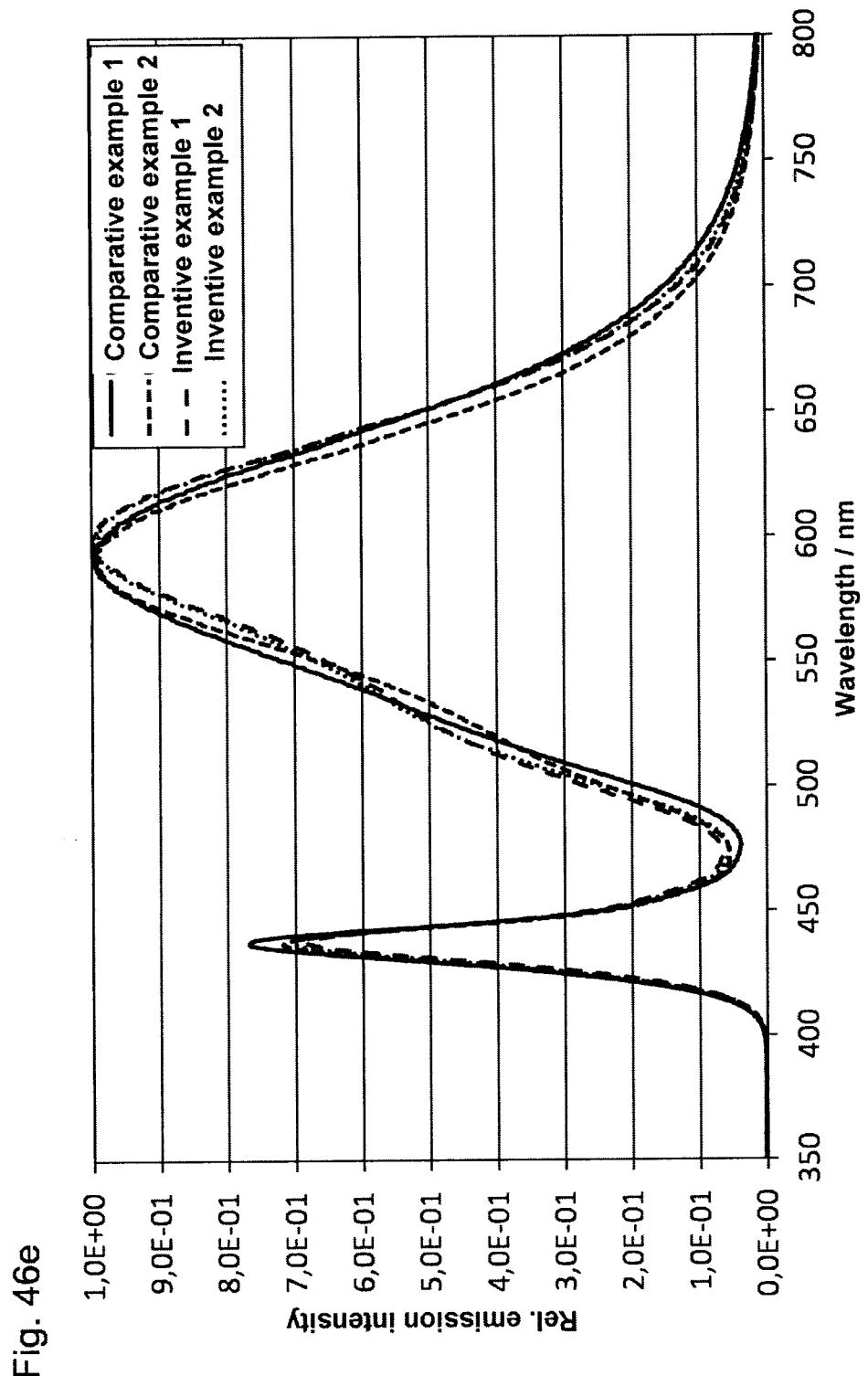


Fig. 47a

Example	CIE-x	CIE-y	CRI	Rel. $\Phi_{\text{v(filled)}}$	Rel. LER
Comp. ex. 1	0,379	0,377	71	100%	100%
Inventive ex. 1	0,380	0,378	70	100,7%	101,3%

Fig. 47b

Example	T (°C)	λ_{dom} LED	CIE-x	CIE-y	CRI	Rel. Φ_{filled} potting/ Φ_{ef} (clear potting)	Rel. LER
Comp. ex. 1	25	444,8	0,372	0,368	71	100%	100%
	85	447,8	0,369	0,370	72	101,6%	100,9%
Inventive ex. 1	25	445,0	0,373	0,370	71	101,3%	101,5%
	85	448,0	0,368	0,371	71	101,3%	101,6%

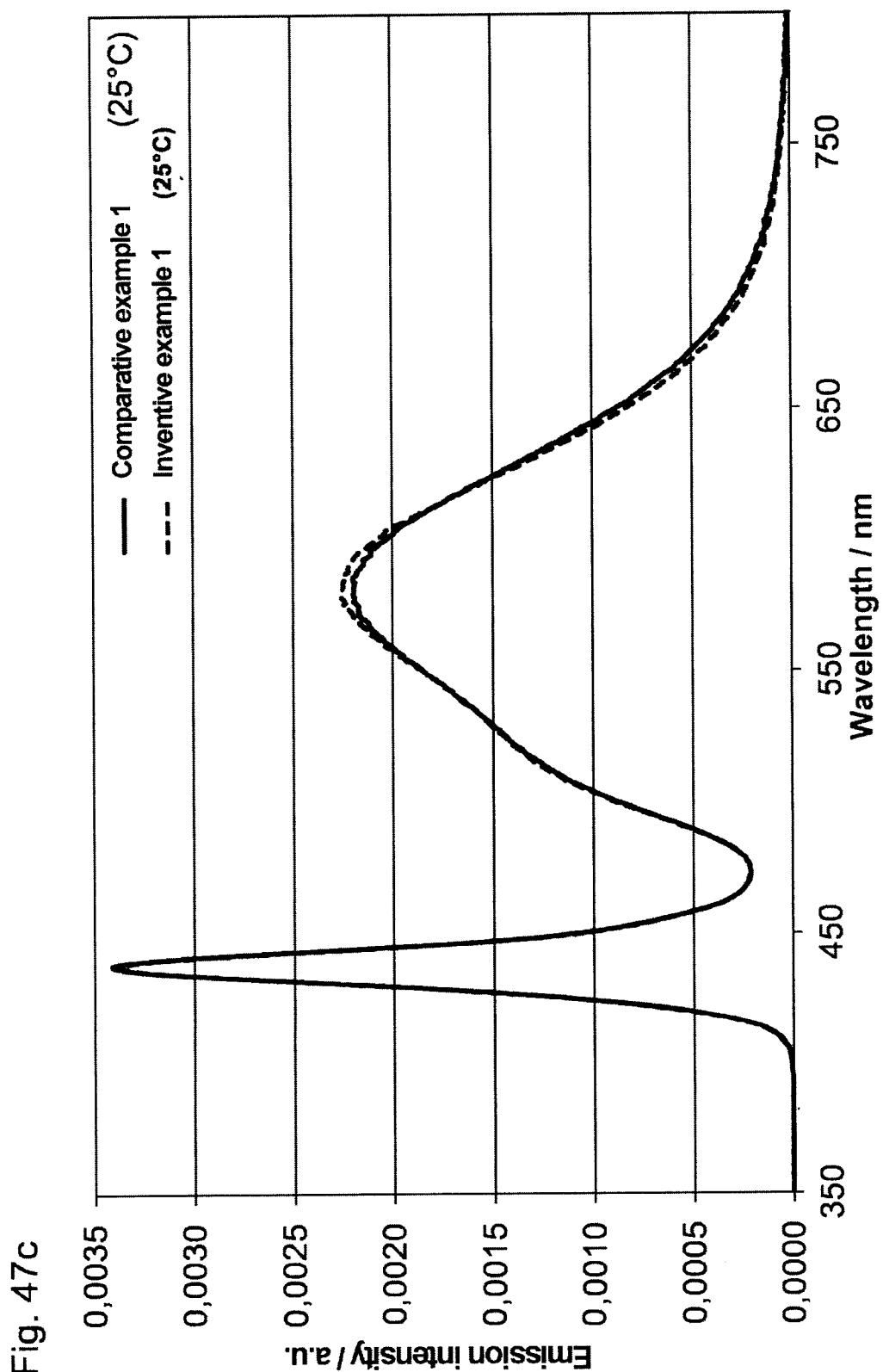
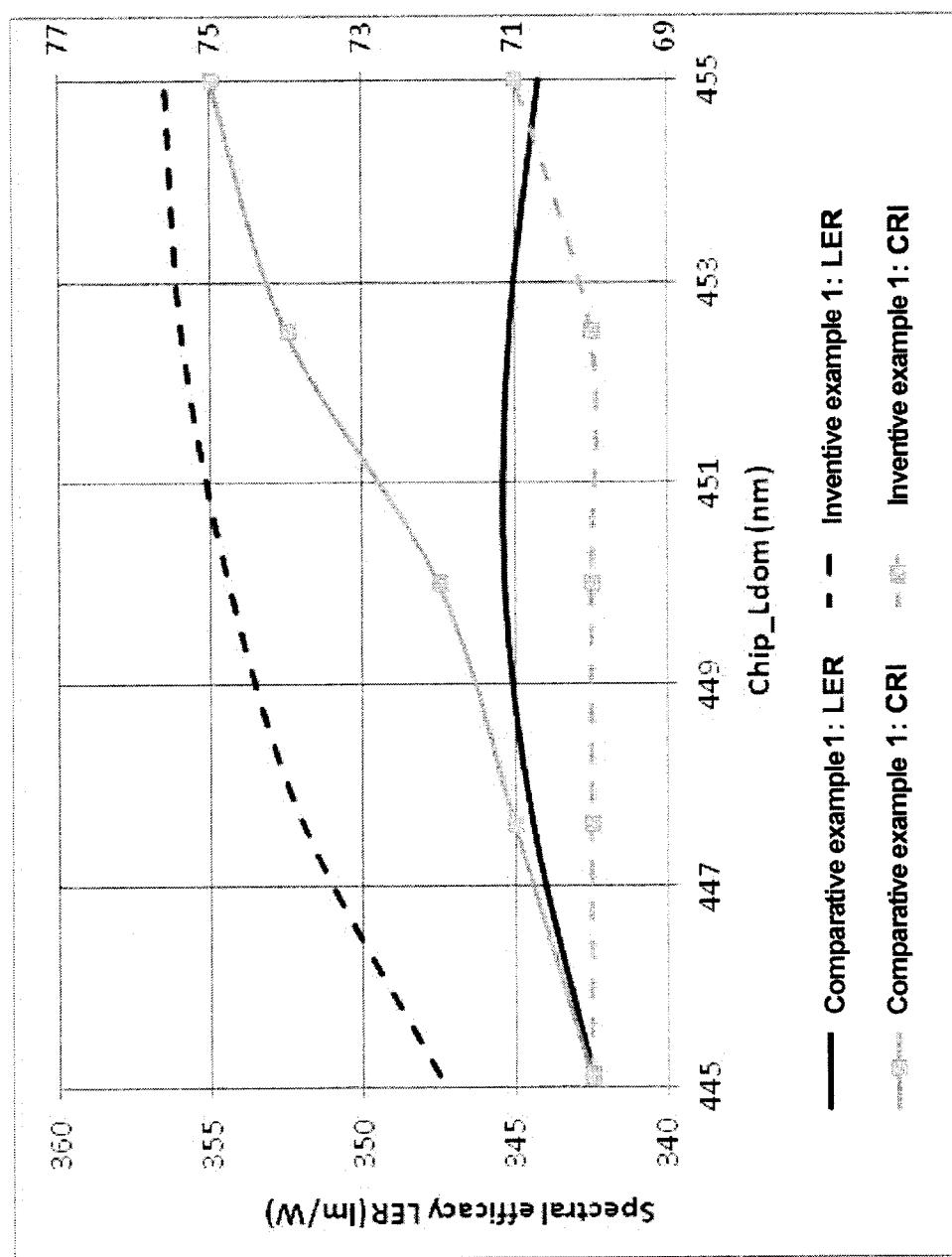


Fig. 47d



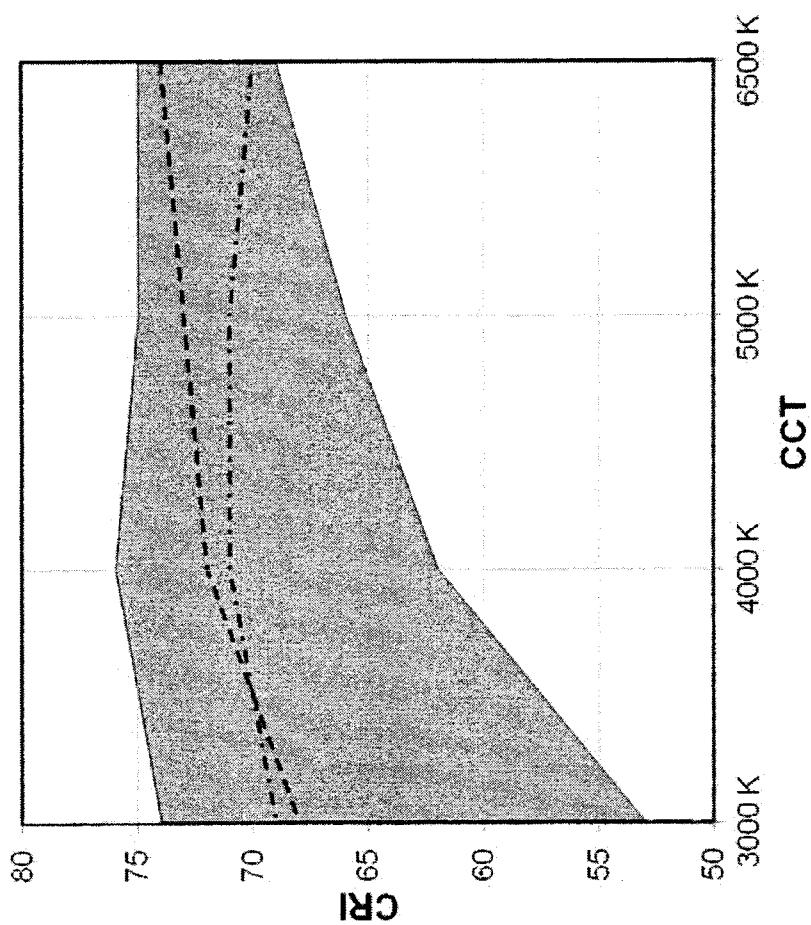


Fig. 47e

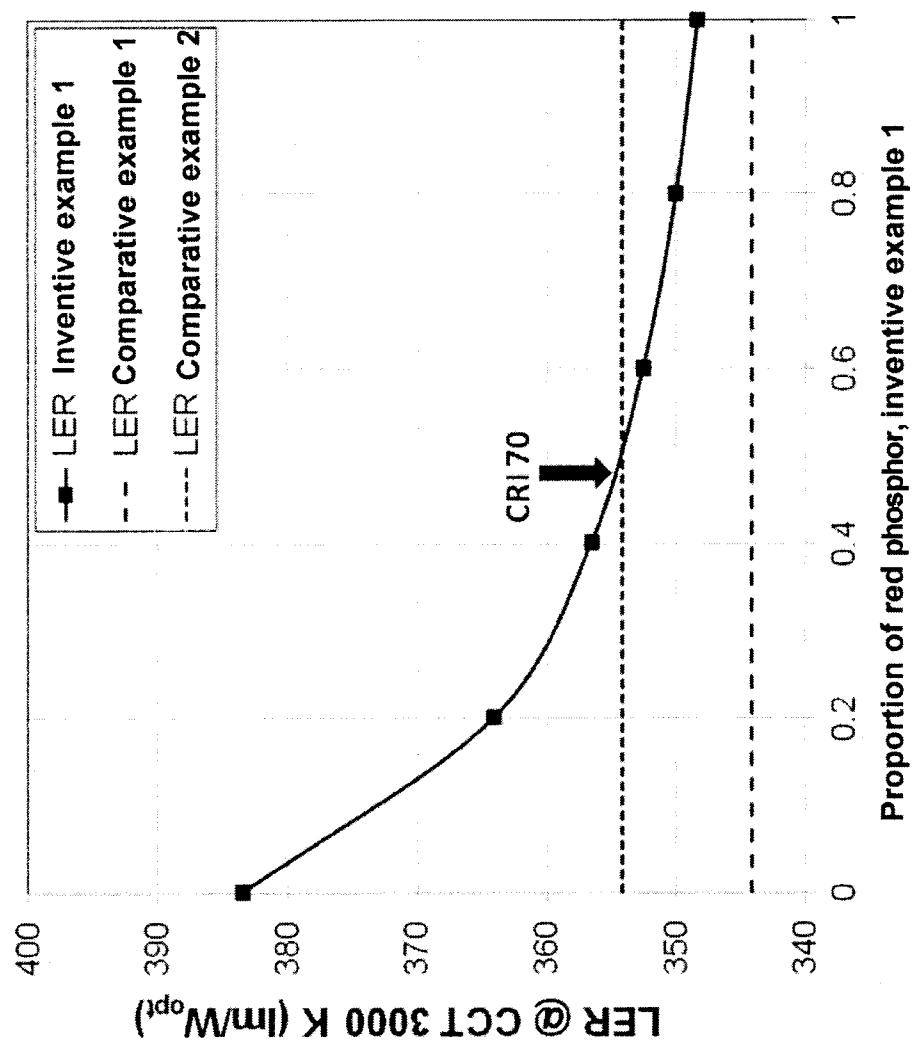


Fig. 47f

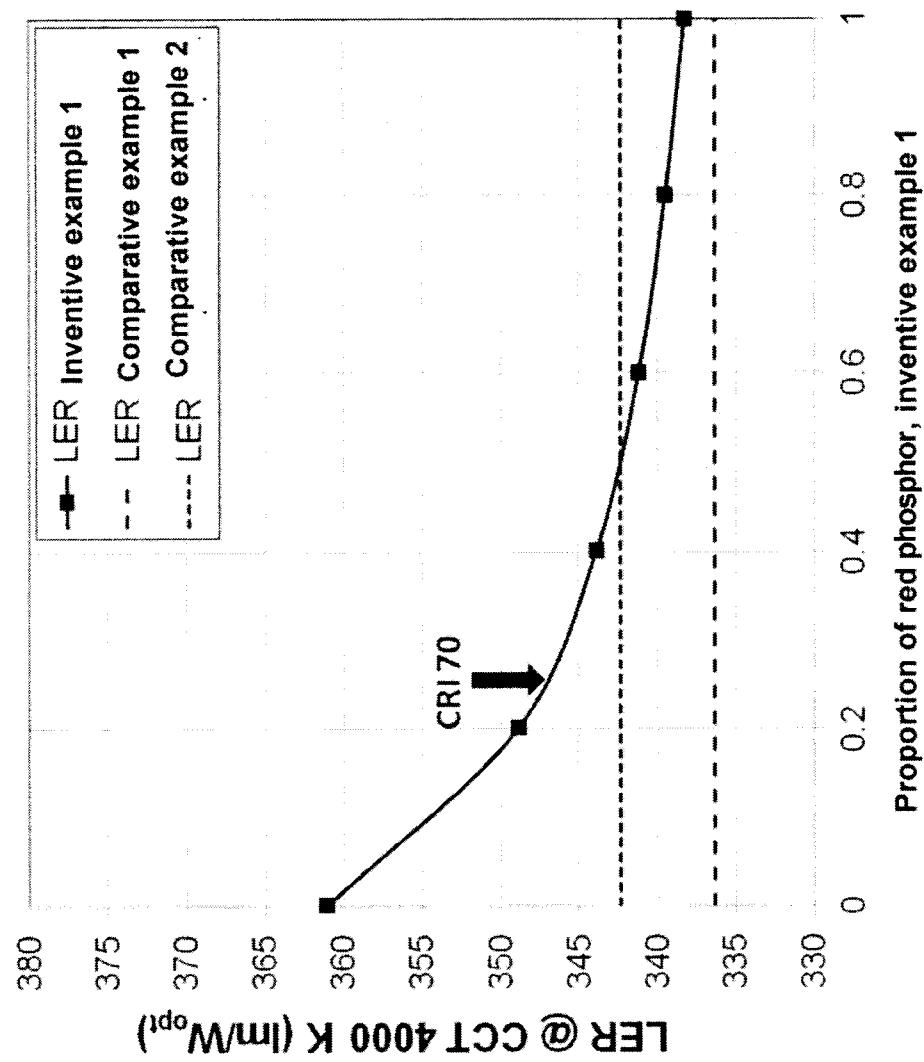


Fig. 47g

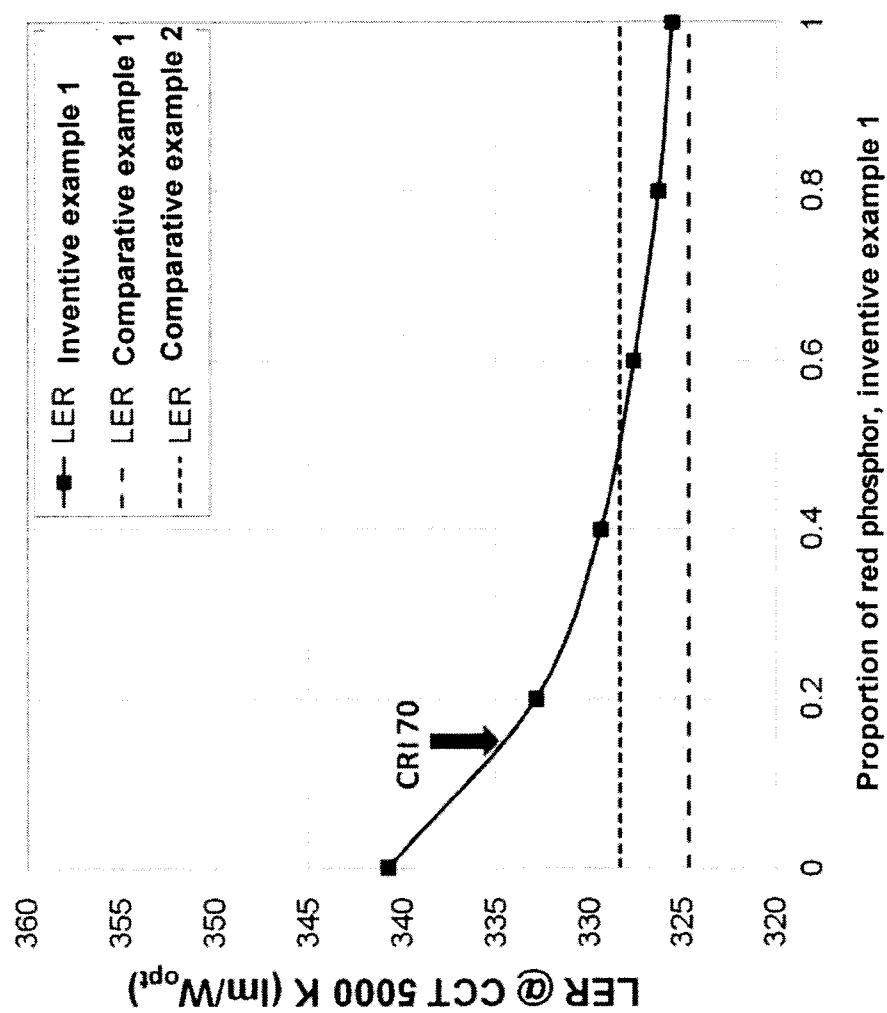


Fig. 47h

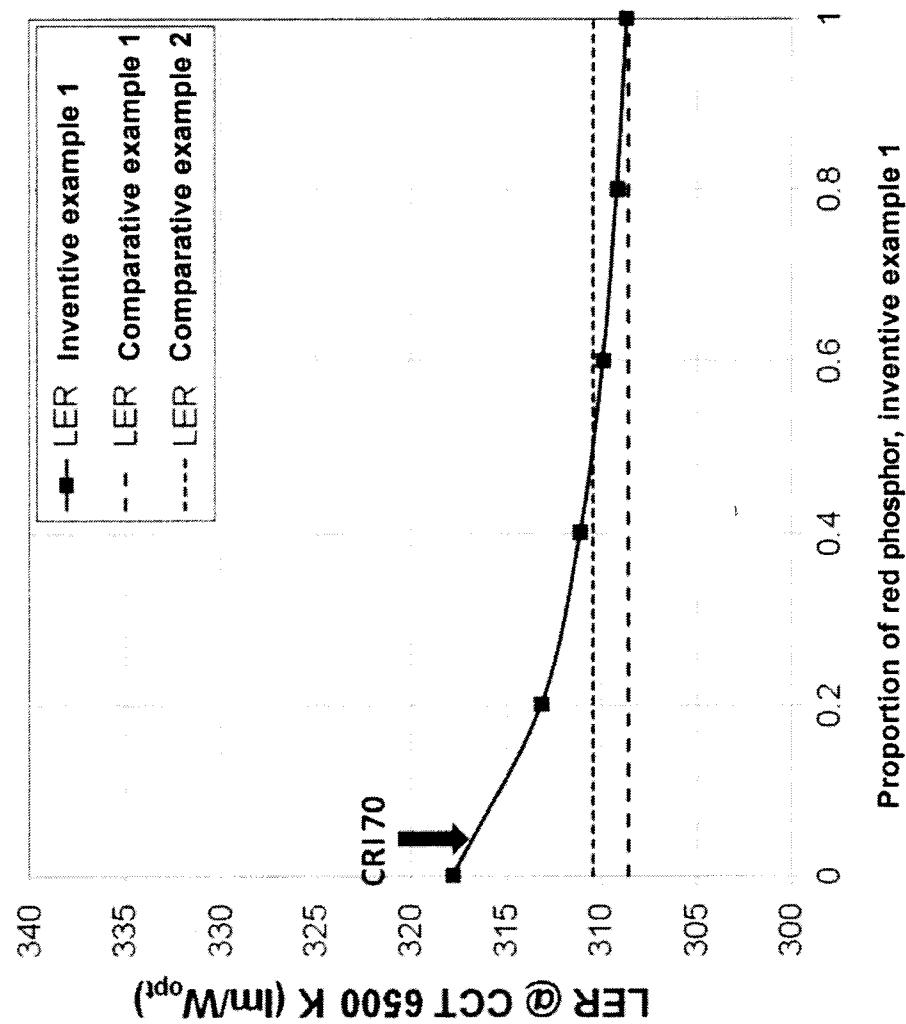


Fig. 47i

Fig. 48a

phosphor	nominal composition	CIE x	CIE y	dominant wavelength / nm	FWHM / nm	LER / lm/W	relative LER	relative ext. QE
comparative example 1	CaAlSiN ₃ :Eu	0,656	0,341	607,9	89	132	100%	100%
embodiment 1	Sr(Sr0.8Ca0.14)Si2Al2N6:Eu (3.7%)	0,659	0,340	608,2	78	172	130%	125%
embodiment 2	Sr(Sr0.86Ca0.14)Si2Al2N6:Eu (10%)	0,675	0,324	613,4	80	127	96%	120%
embodiment 3	Sr(Sr0.86Ca0.14)Si2Al2N6:Eu (20%)	0,687	0,311	618,8	85	88	66%	99%

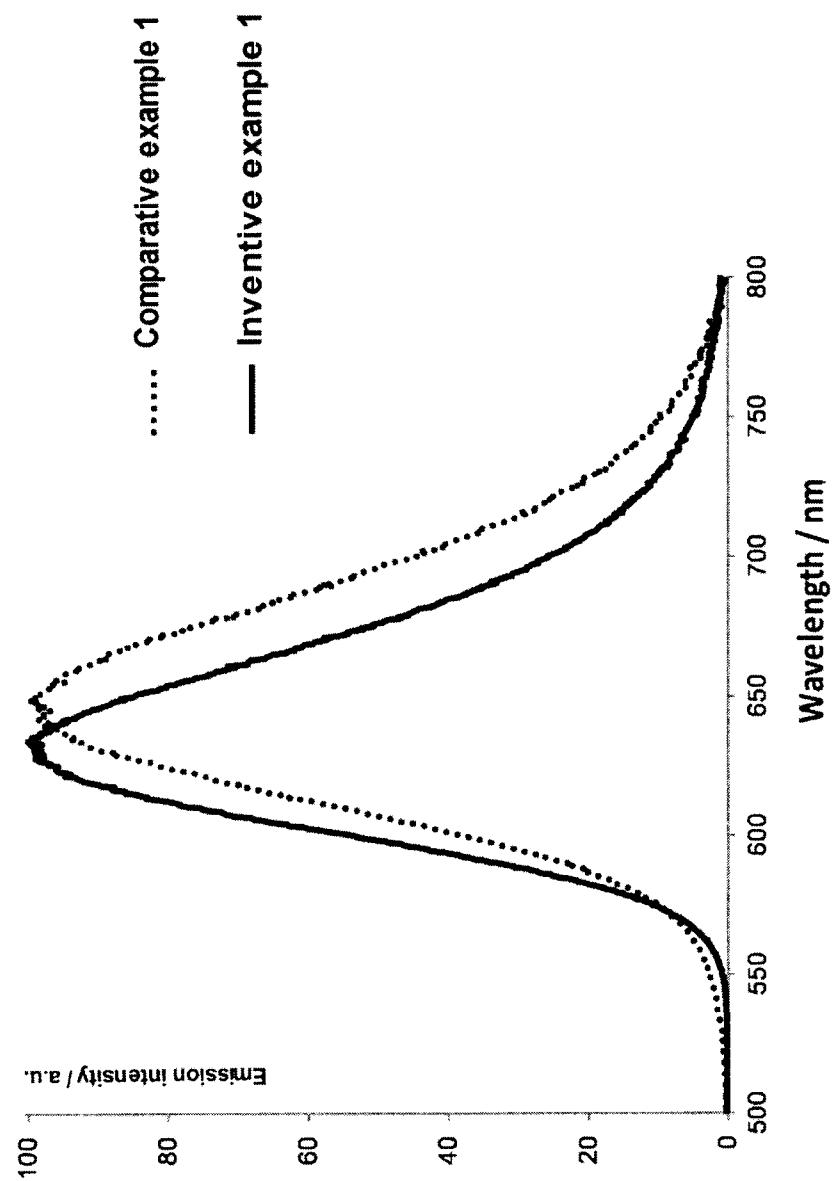


Fig. 48b

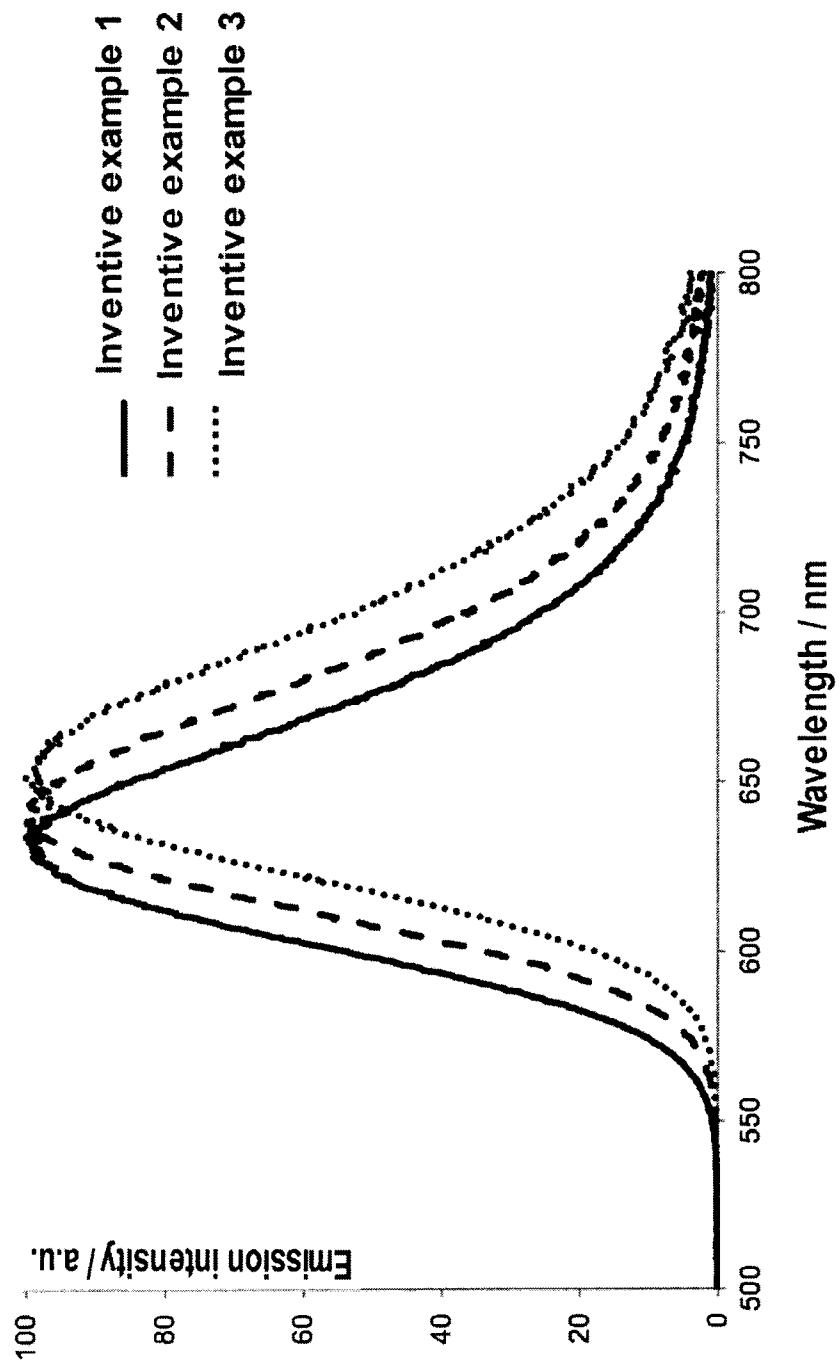


Fig. 48c

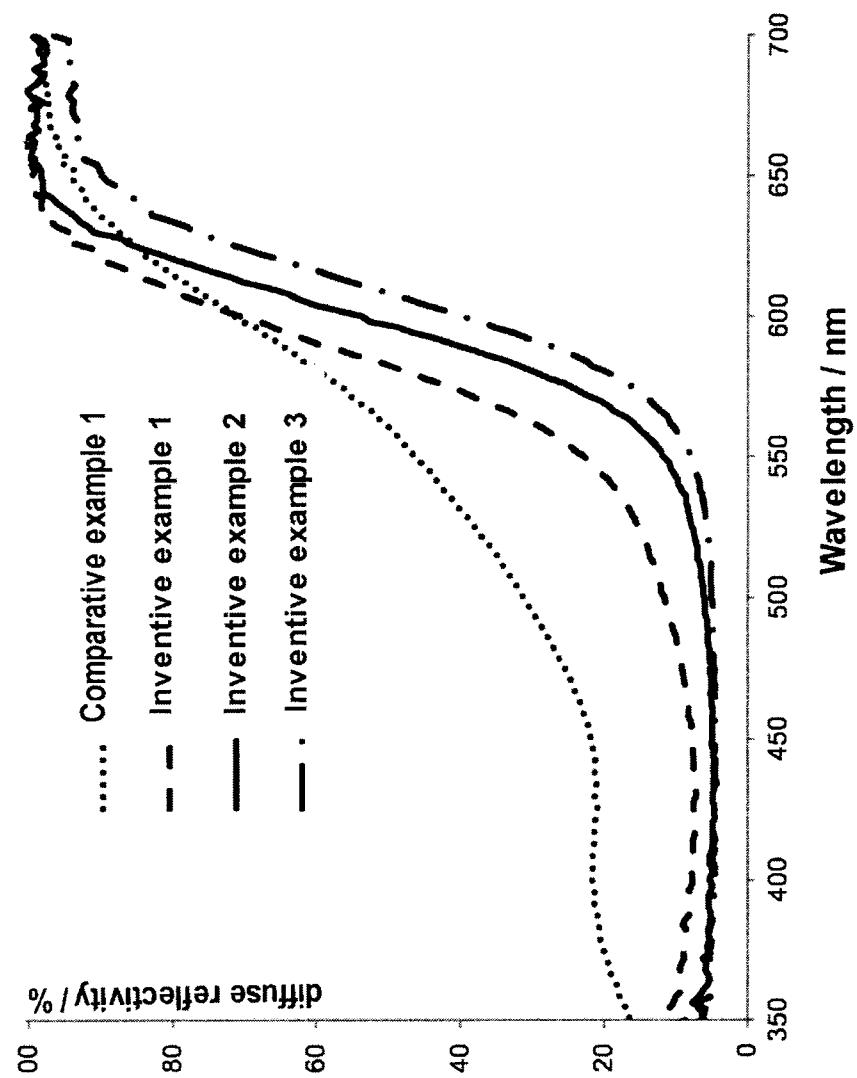


Fig. 48d

Fig. 48e

solution	comparative example LED1	embodiment LED1
LED dominant wavelength / nm	452,3	452,3
green phosphor	LuAG:Ce	LuAG:Ce
red phosphor	CaAlSiN ₃ :Eu	Sr(Sr _{0.8} Ca _{0.2})Si ₂ Al ₂ N ₆ :Eu (3.7%)
total phosphor concentration	11,8%	13,0%
green:red ratio	6,3:1	31:1
CIE x/y (unfiltered)	0,286 / 0,273	0,286 / 0,272
CIE u'v' (unfiltered)	0,200 / 0,431	0,201 / 0,430
CIE x/y (red channel, filtered)	0,653 / 0,330	0,650 / 0,333
CIE x/y (green channel, filtered)	0,292 / 0,590	0,295 / 0,587
CIE x/y (blue channel, filtered)	0,149 / 0,063	0,149 / 0,062
CIE u'/v' (red channel, filtered)	0,463 / 0,525	0,456 / 0,526
CIE u'/v' (green channel, filtered)	0,123 / 0,559	0,125 / 0,559
CIE u'/v' (blue channel, filtered)	0,173 / 0,164	0,173 / 0,163
color gamut overlap vs. sRGB / rec. 709 in x/y color space	98,2%	98,0%
color gamut overlap vs. sRGB / rec. 709 in u'v' color space	98,7%	98,3%
relative conversion efficiency	100%	105%
relative LER	100%	104%

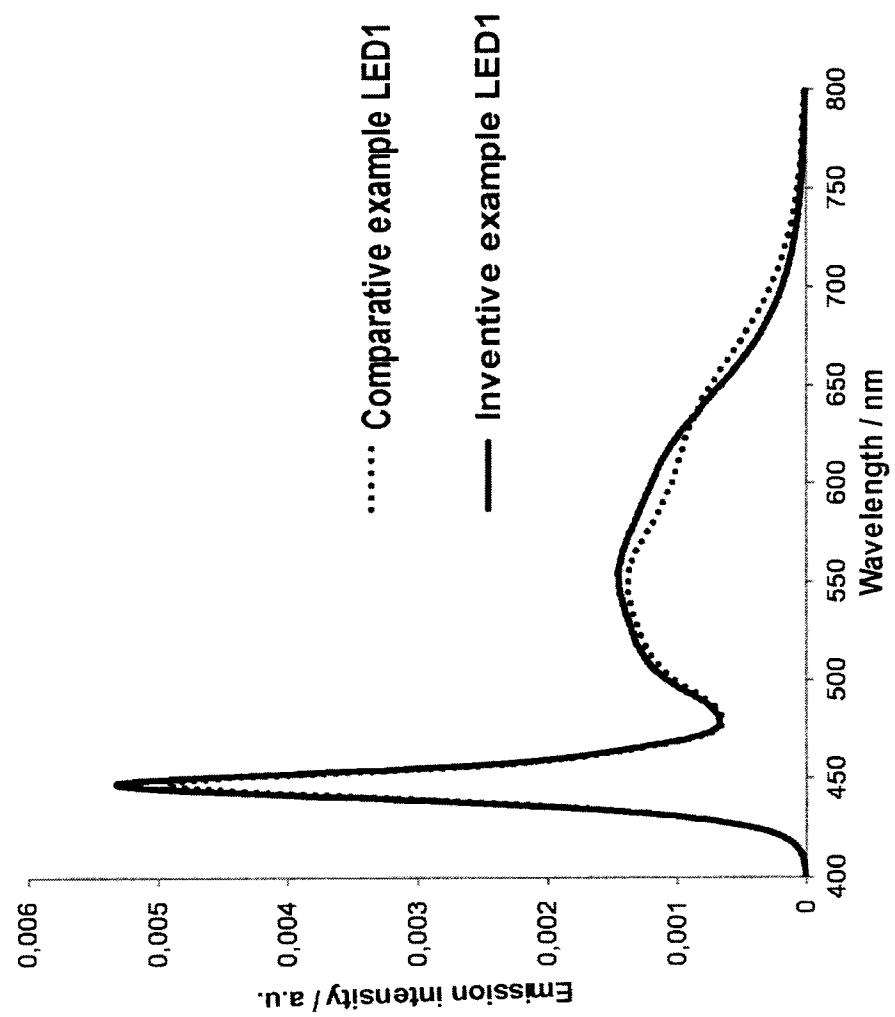


Fig. 48f

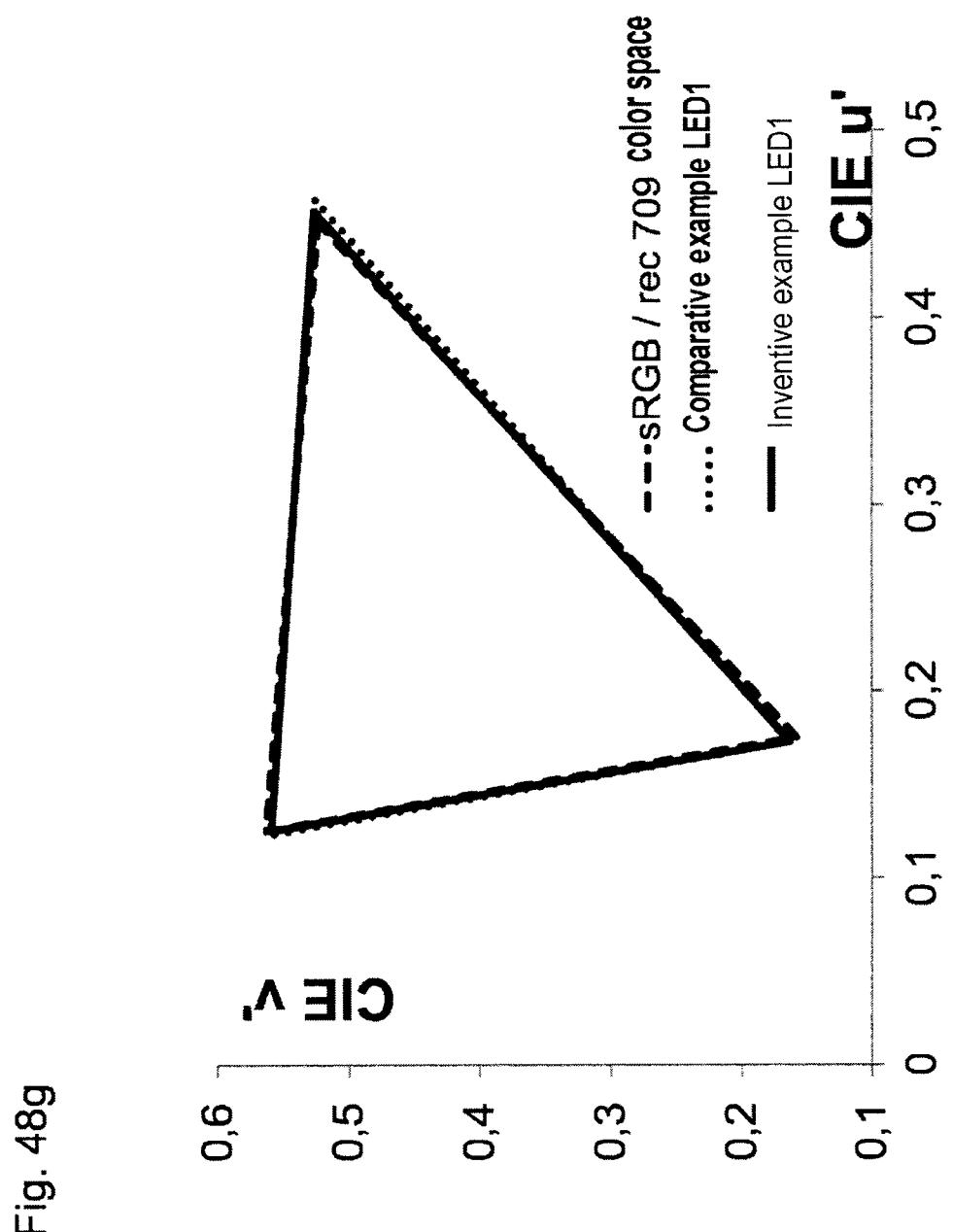


Fig. 48g

Fig. 48h

solution	comparative example LED2	embodiment LED2	embodiment LED3
LED dominant wavelength / nm	450,6	446,9	450,7
green phosphor	Nitrido-Orthosilicate:Eu	Nitrido-Orthosilicate:Eu	Nitrido-Orthosilicate:Eu
red phosphor	CaAlSiN3:Eu	Sr(Sr0.86Ca0.14)Si2Al2N6:Eu (10%)	Sr(Sr0.86Ca0.14)Si2Al2N6:Eu (20%)
total phosphor concentration	9,3%	8,8%	11,0%
green:red ratio	2,2:1	13,5:1	12,2:1
CIE x/y (unfiltered)	0,274 / 0,247	0,273 / 0,249	0,274 / 0,249
CIE u'v' (unfiltered)	0,202 / 0,410	0,200 / 0,412	0,201 / 0,412
CIE x/y (red channel, filtered)	0,663 / 0,319	0,657 / 0,319	0,666 / 0,316
CIE x/y (green channel, filtered)	0,271 / 0,604	0,271 / 0,613	0,265 / 0,604
CIE x/y (blue channel, filtered)	0,151 / 0,057	0,154 / 0,057	0,150 / 0,060
CIE u'v' (red channel, filtered)	0,483 / 0,522	0,477 / 0,521	0,488 / 0,521
CIE u'v' (green channel, filtered)	0,112 / 0,560	0,111 / 0,562	0,109 / 0,559
CIE u'v' (blue channel, filtered)	0,178 / 0,153	0,182 / 0,151	0,175 / 0,158
color gamut overlap vs. DCI-P3 in x/y color space	81,6%	81,7%	82,7%
color gamut overlap vs. DCI-P3 in u'v' color space	90,5%	89,4%	91,7%

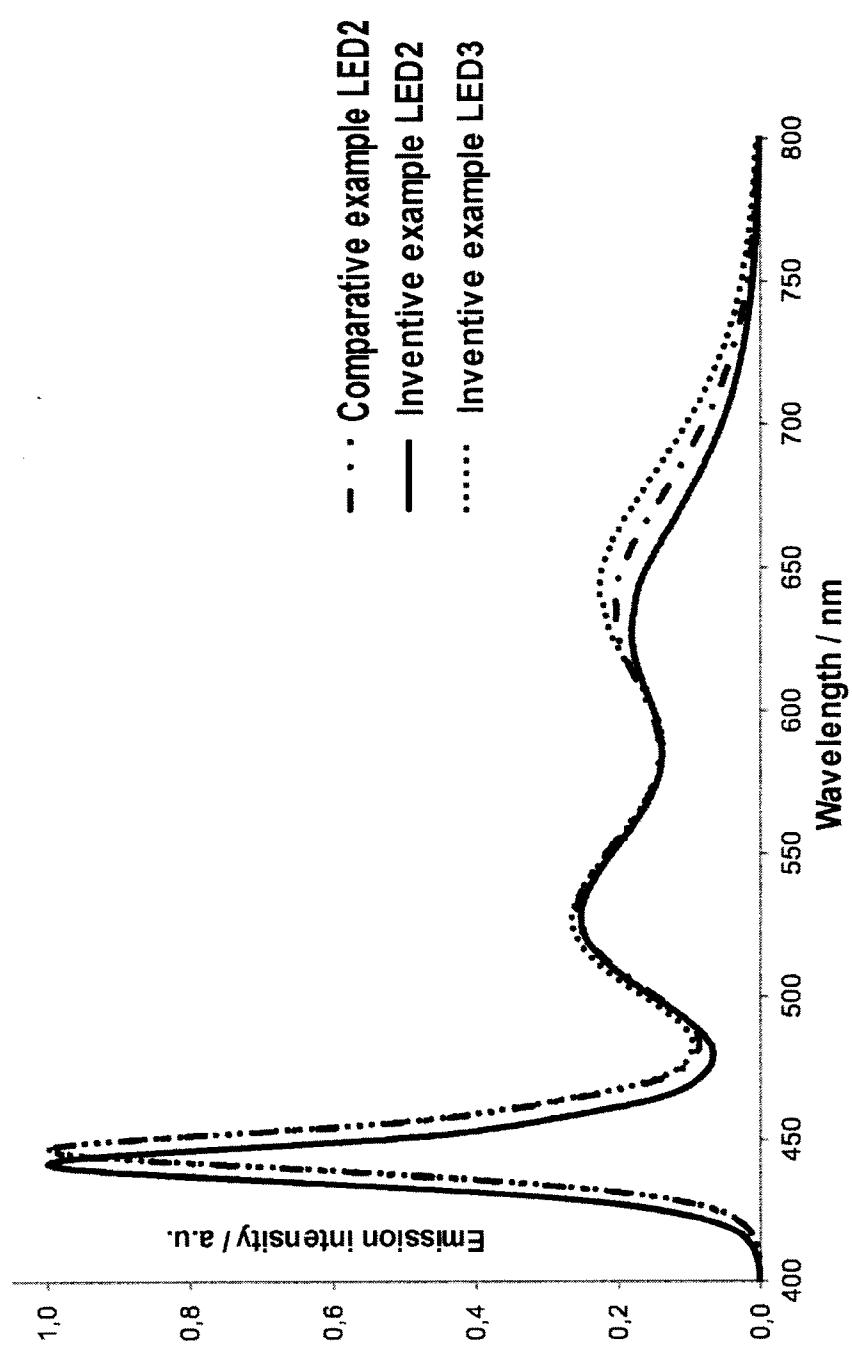
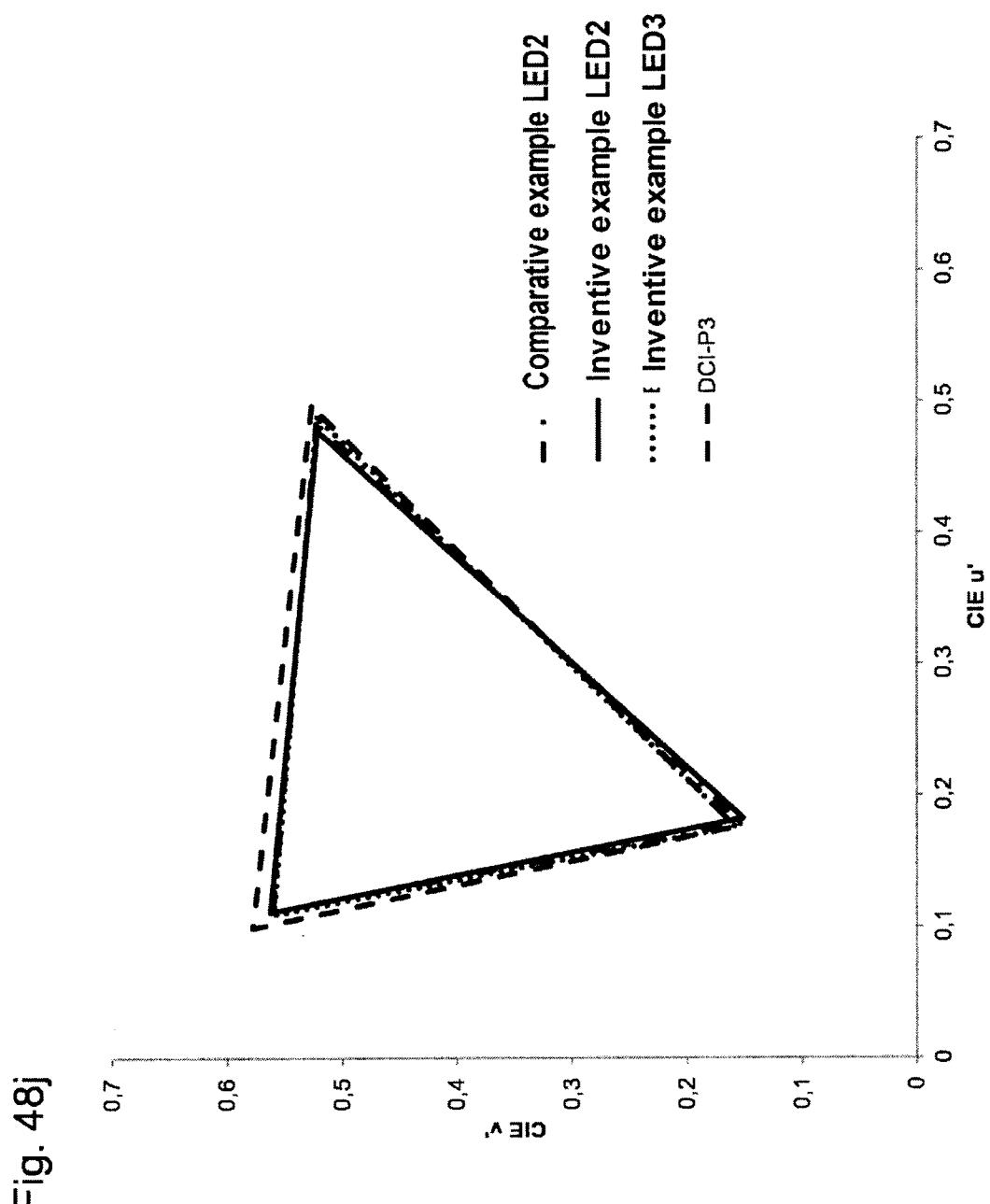


Fig. 48i



Current/mA	CIE-x (Comp. ex. 1)	CIE-y (Comp. ex. 1)	CIE-x (Inventive ex. 1)	CIE-y (Inventive ex. 1)	$\Phi_{\text{v},\text{filled}}$ potting)	$\Phi_{\text{v},\text{clear}}$ potting)	Relative LER (luminous efficiency)
40	0,346	0,350	0,346	0,348	101,2%	101,6%	
100	0,346	0,351	0,346	0,349	101,3%	101,5%	
200	0,346	0,351	0,346	0,350	101,3%	101,5%	
350	0,345	0,352	0,346	0,350	101,4%	101,4%	
500	0,344	0,352	0,345	0,350	101,6%	101,3%	
700	0,343	0,352	0,345	0,350	101,9%	101,3%	
1000	0,342	0,351	0,344	0,350	102,1%	101,2%	

Fig. 49a

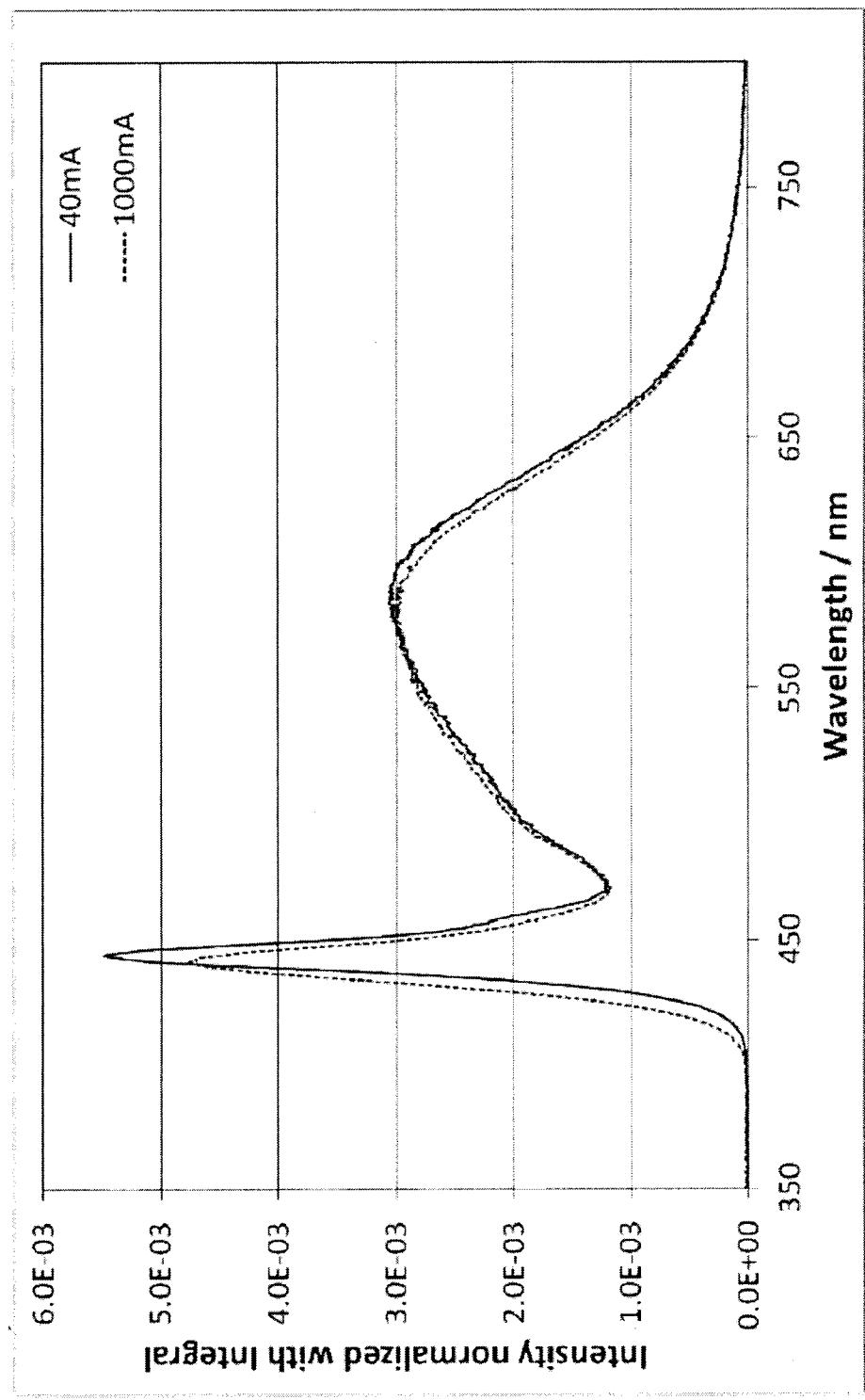


Fig. 49b

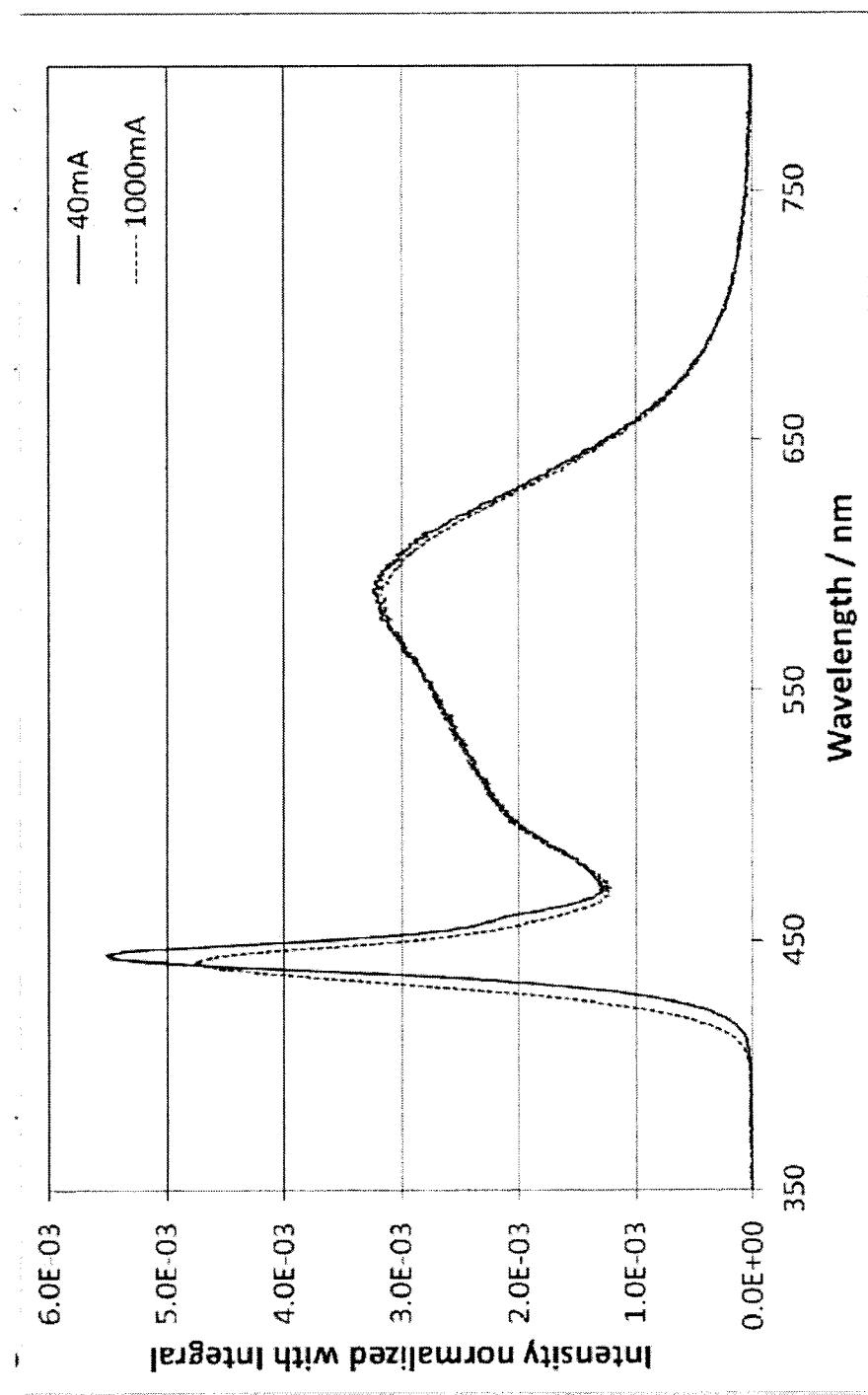


Fig. 49c

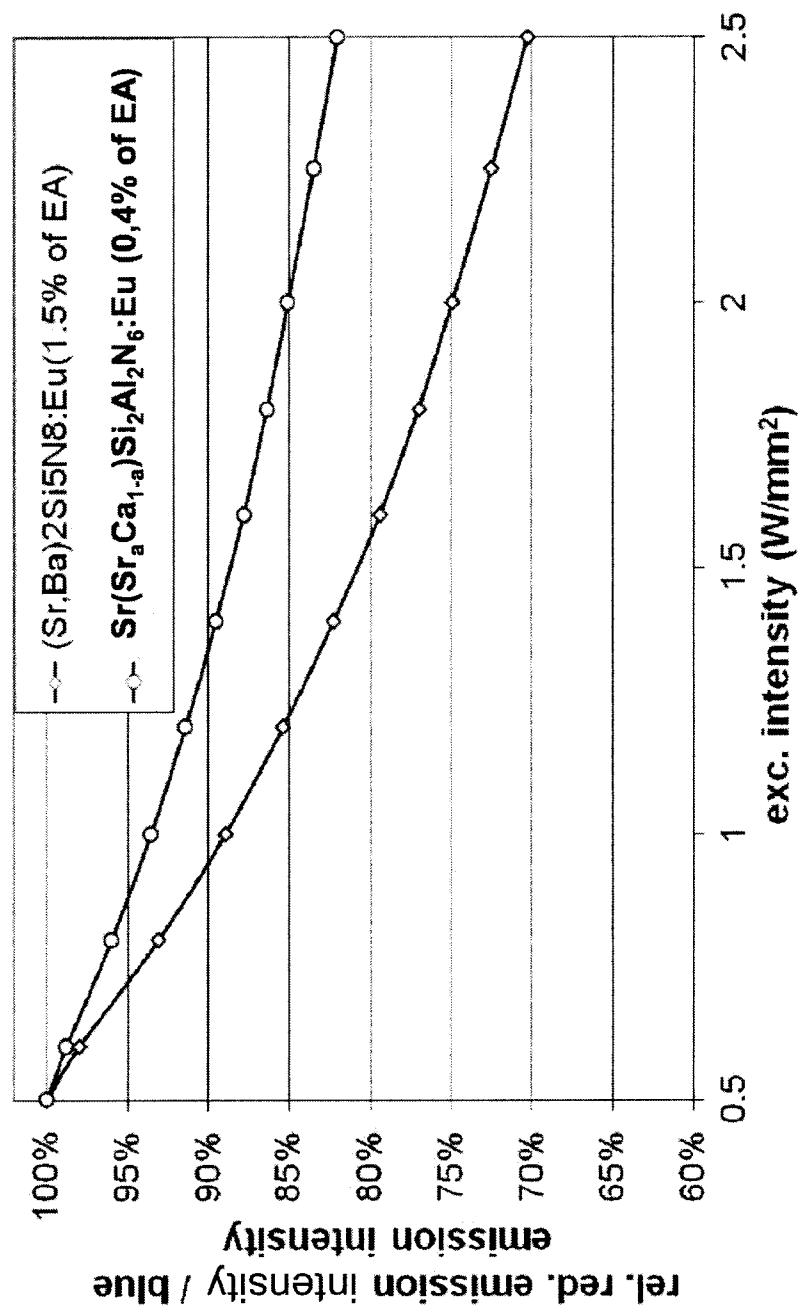
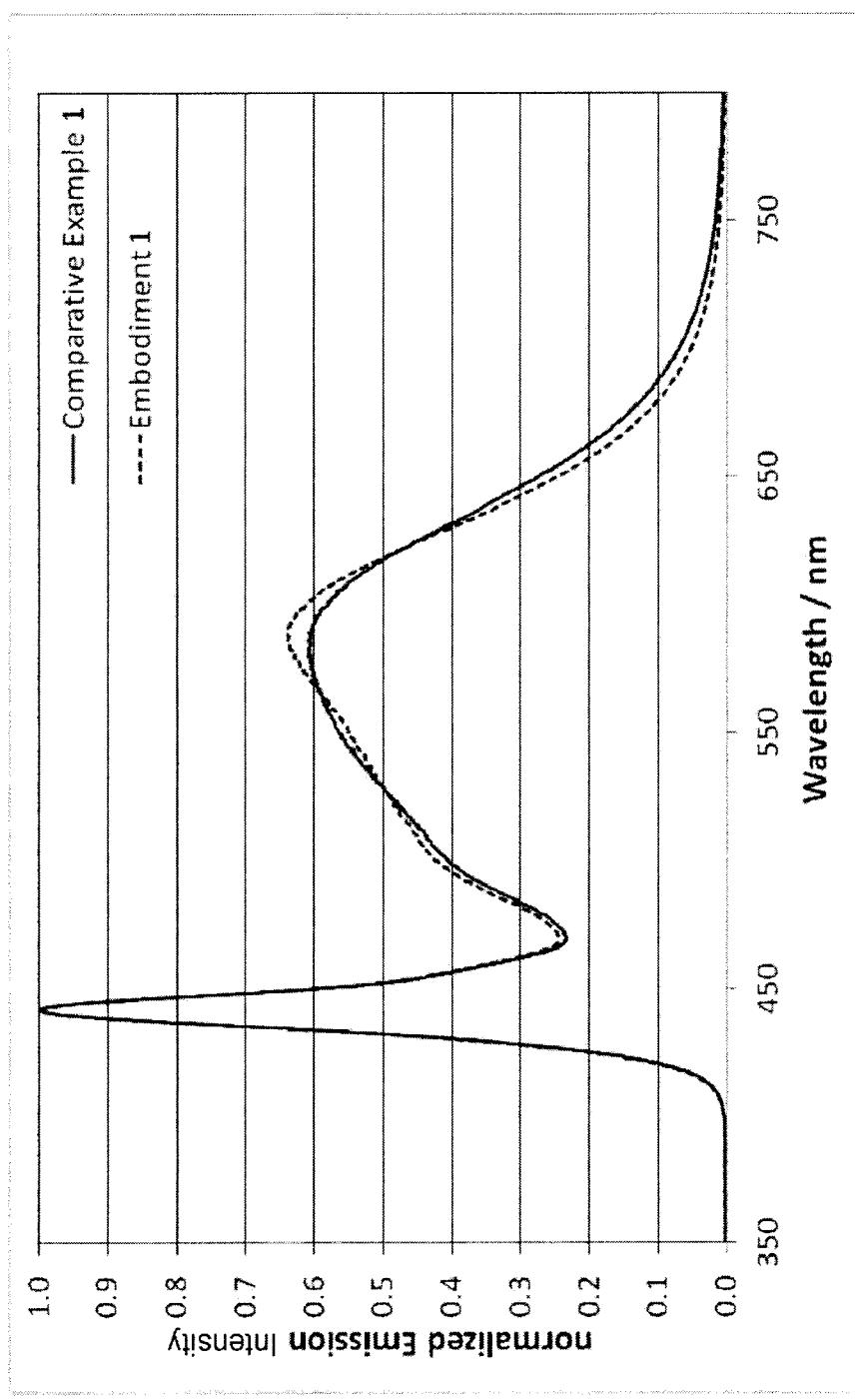


Fig. 49d

Fig. 49e



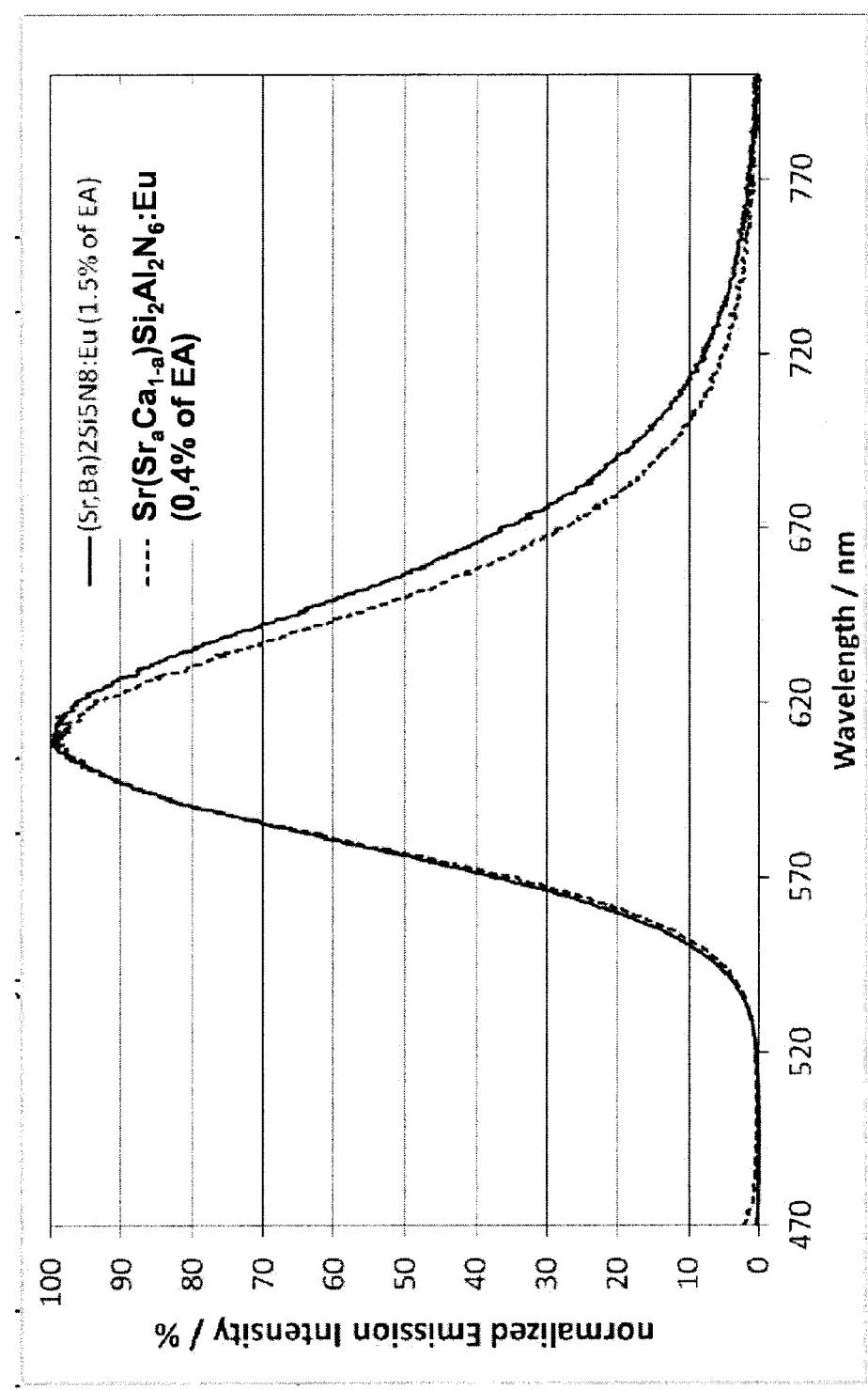


Fig. 49f

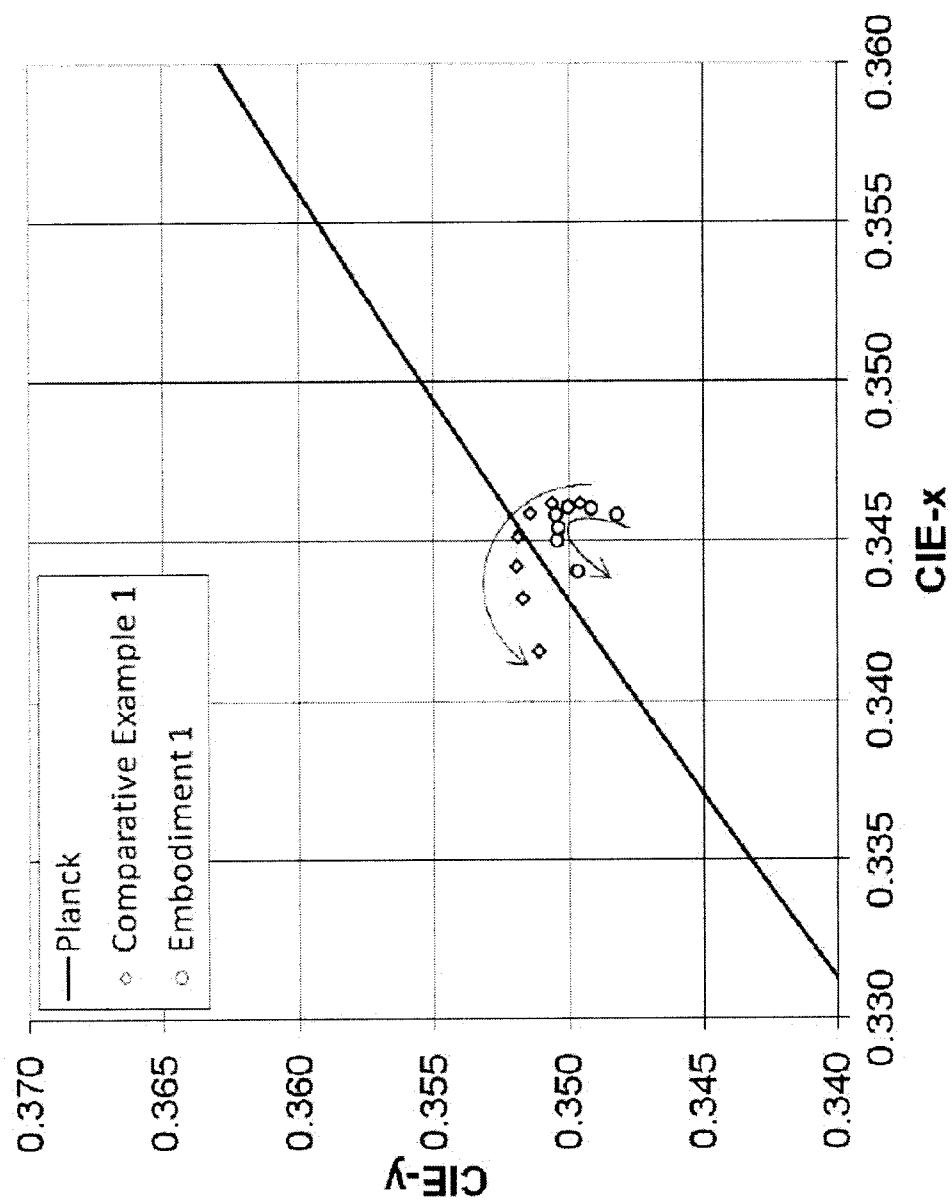


Fig. 49g

CurrentmA	CIE-x (Comp. ex. 2)	CIE-y (Comp. ex. 2)	CIE-x (Inventive ex. 2)	CIE-y (Inventive ex. 2)	Relative LER		(luminous efficiency)
					Φ_{filled}	$\Phi_{\text{clear}}/\Phi_{\text{filled}}$	
40	0,563	0,430	0,562	0,429	116,2%	110,4%	
100	0,563	0,430	0,562	0,428	116,1%	110,3%	
200	0,562	0,430	0,561	0,428	116,5%	110,1%	
350	0,561	0,430	0,561	0,428	117,1%	109,7%	
500	0,561	0,431	0,561	0,429	118,3%	109,5%	
700	0,560	0,431	0,560	0,429	118,9%	109,1%	
1000	0,559	0,432	0,559	0,429	117,0%	108,5%	

Fig. 50a

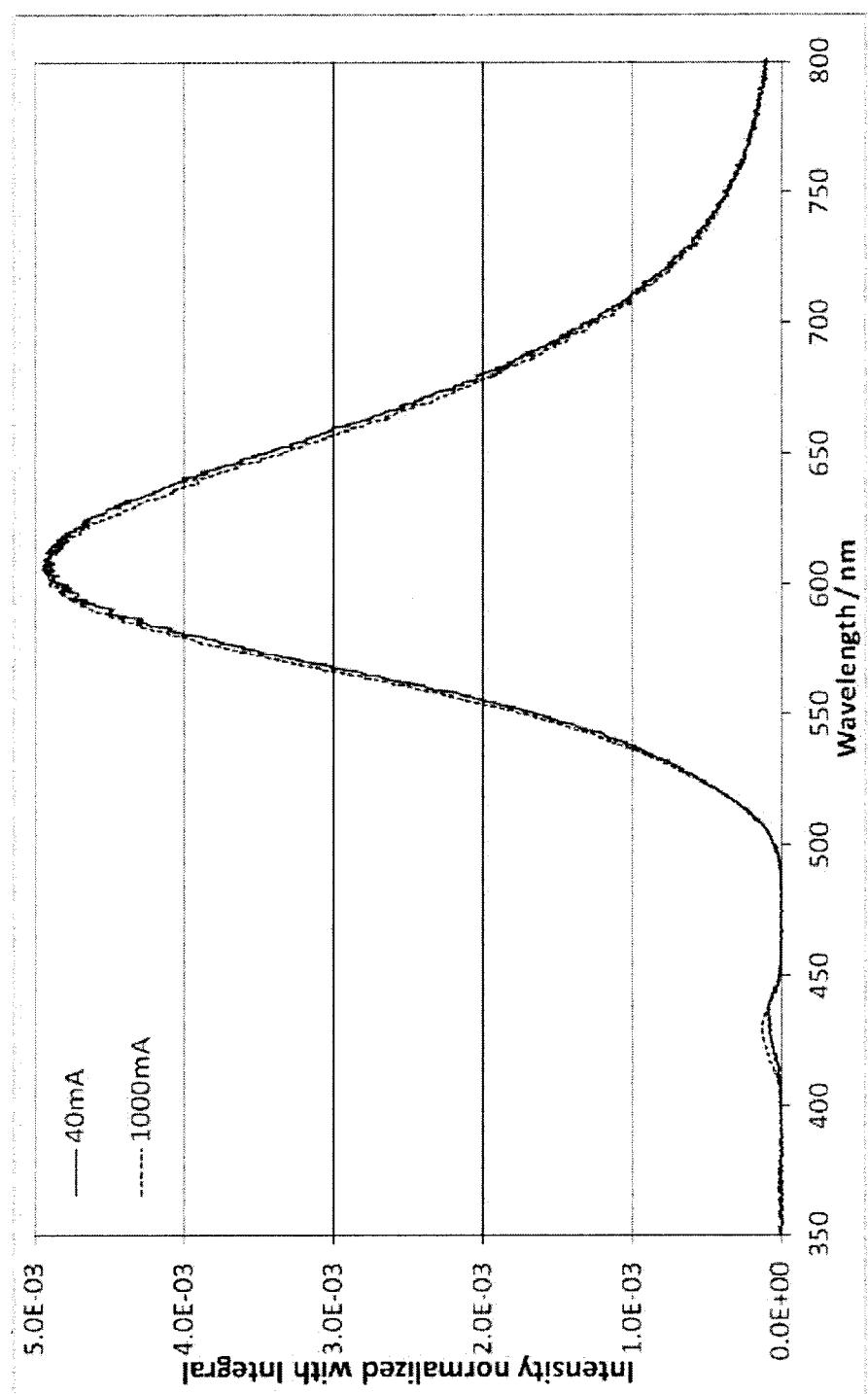


Fig. 50b

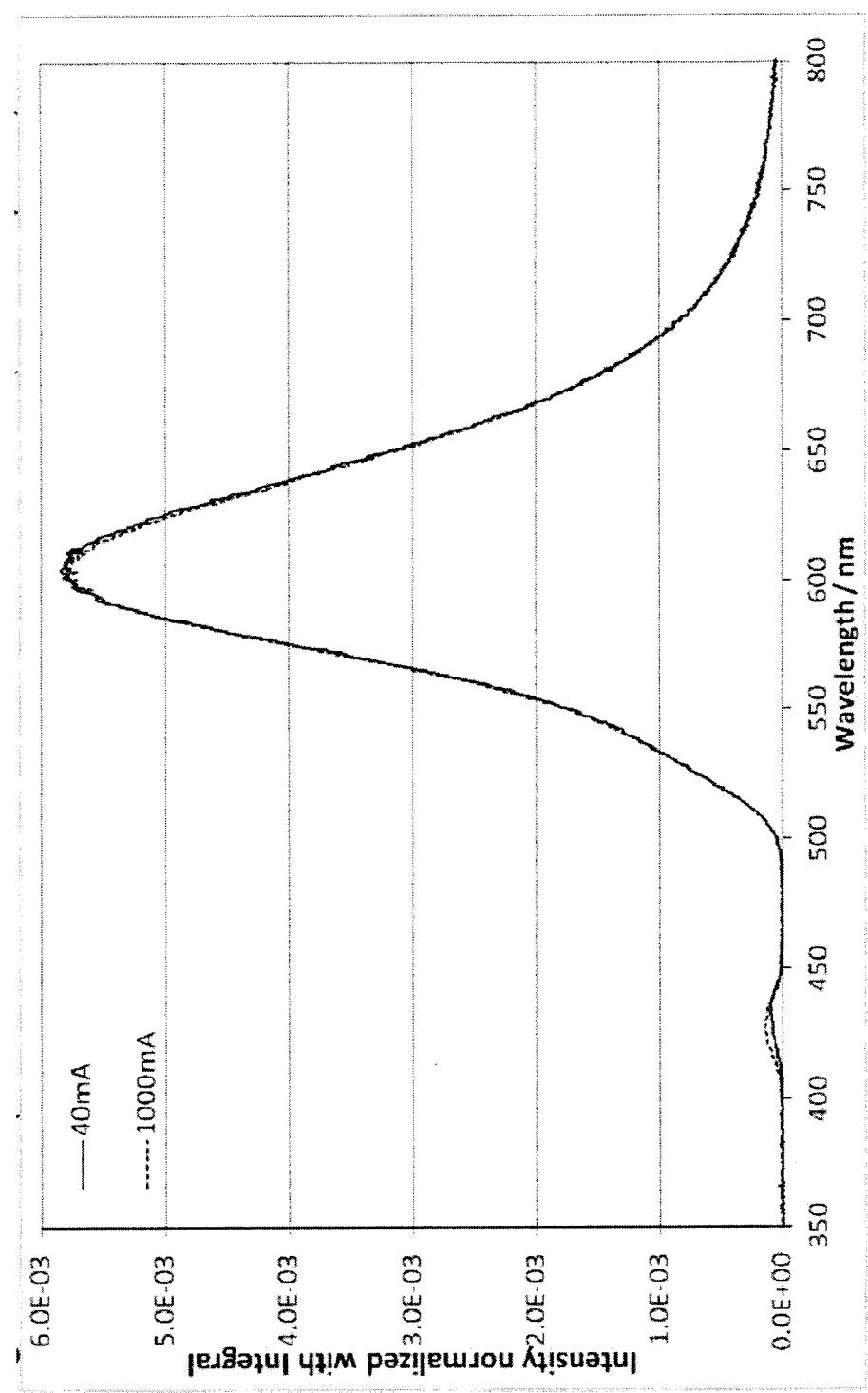


Fig. 50c

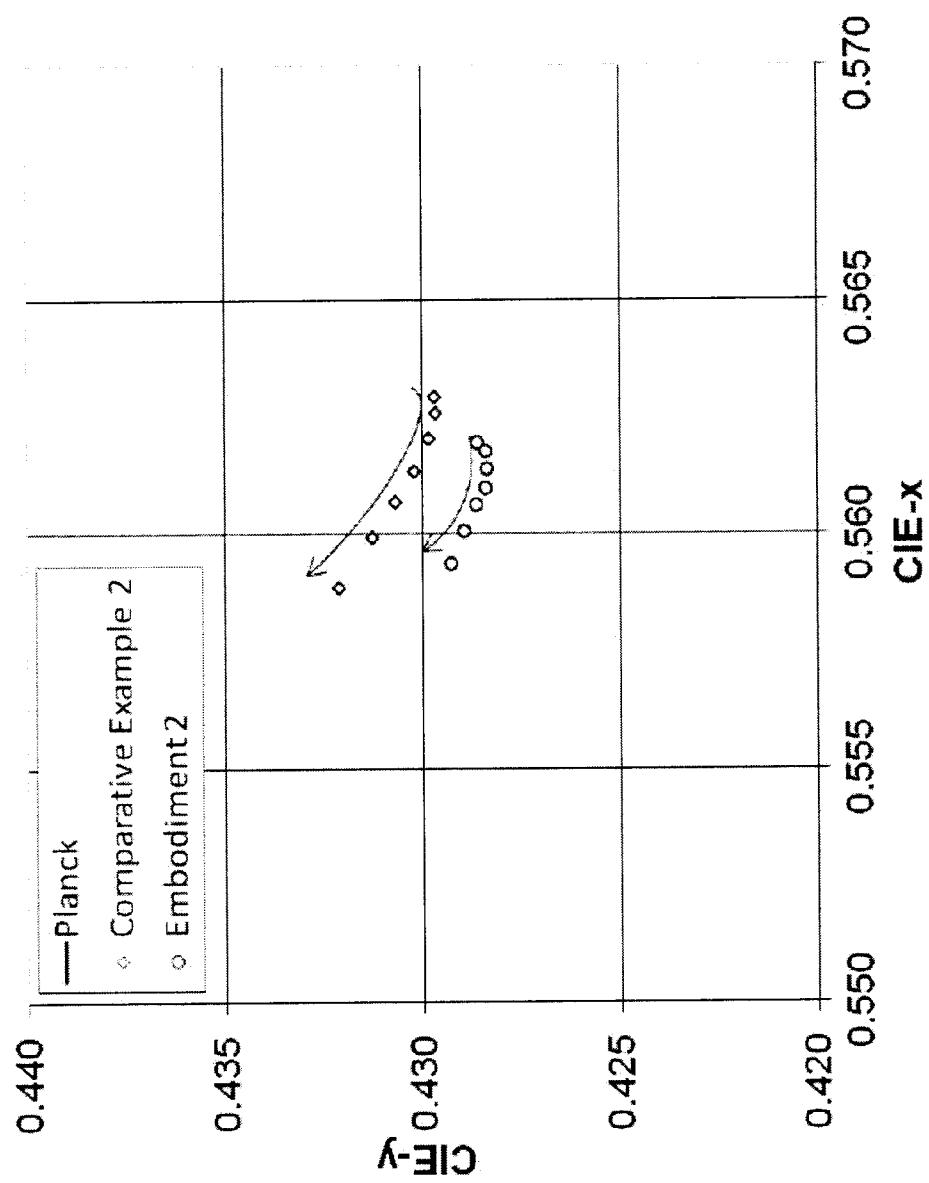


Fig. 50d

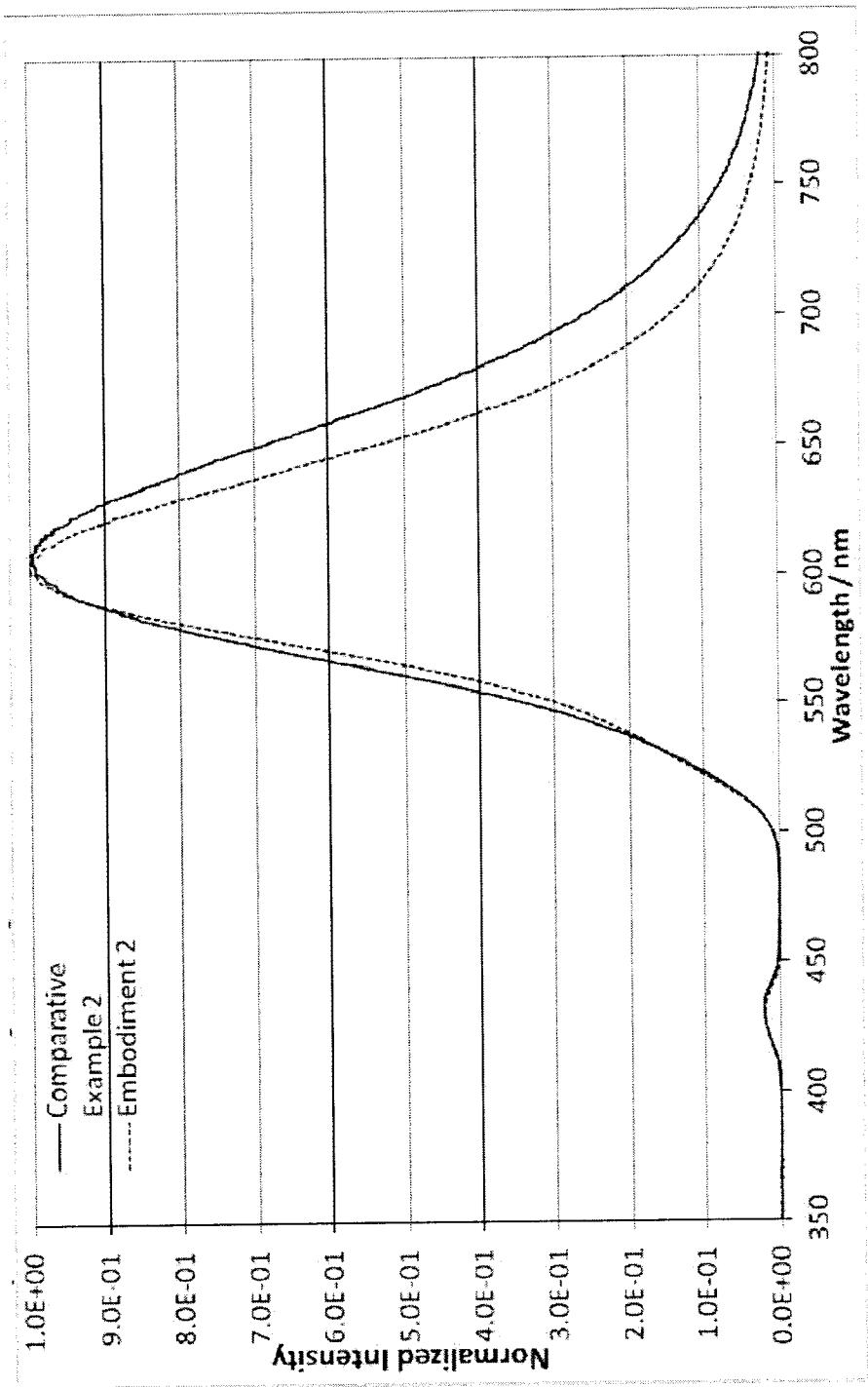


Fig. 50e

Fig. 51a

Example	$\lambda_{\text{down/blue}}$ LED	Composition of green phosphor	Composition of red phosphor	Total phos- phor conc. (% by wt.)	Green/red ratio	Abs. green conc. (% by wt.)	Abs. red conc. (% by wt.)	CCT/K	CRI	R9	Rel. conversion efficiency
Comp. ex. 1	446	$\text{Y}_3(\text{Al}_{0.6}\text{Ga}_{0.4})_5\text{O}_{12}$ Ce (ca 2% Ce)	$\text{Sr}_{0.7}\text{Ca}_{0.3}\text{AlSiN}_3$:Eu (0,5% Eu)	15,0%	2,0	10,0%	5,0%	2713	81	9	101%
Comp. ex. 2	446	$\text{Lu}_3(\text{Al}_{0.75}\text{Ga}_{0.25})_5\text{O}_{12}$ Ce (ca 2,5 % Ce)	$(\text{Ca}_{0.1}\text{Sr}_{0.4}\text{Ba}_{0.5})_2\text{Si}_5\text{N}_8$:Eu (3,25% Eu)	16,7%	4,6	13,7%	3,0%	2680	80	11	100%
Intentional ex. 1	446	$\text{Y}_3(\text{Al}_{0.6}\text{Ga}_{0.4})_5\text{O}_{12}$ Ce (ca 2% Ce)	$\text{Sr}(\text{Sr}_{0.86}\text{Ca}_{0.14})\text{Si}_2\text{Al}_2\text{N}_6$:Eu (3% Eu)	15,7%	7,55	13,9%	1,8%	2708	83	10	107%
Intentional ex. 2	446	$\text{Lu}_3\text{Al}_5\text{O}_{12}$: Ce (ca 1,5 % Ce)	$\text{Sr}(\text{Sr}_{0.80}\text{Ca}_{0.20})\text{Si}_2\text{Al}_2\text{N}_6$:Eu (2% Eu)	27,5%	12,4	25,4%	2,1%	2685	81	9	110%

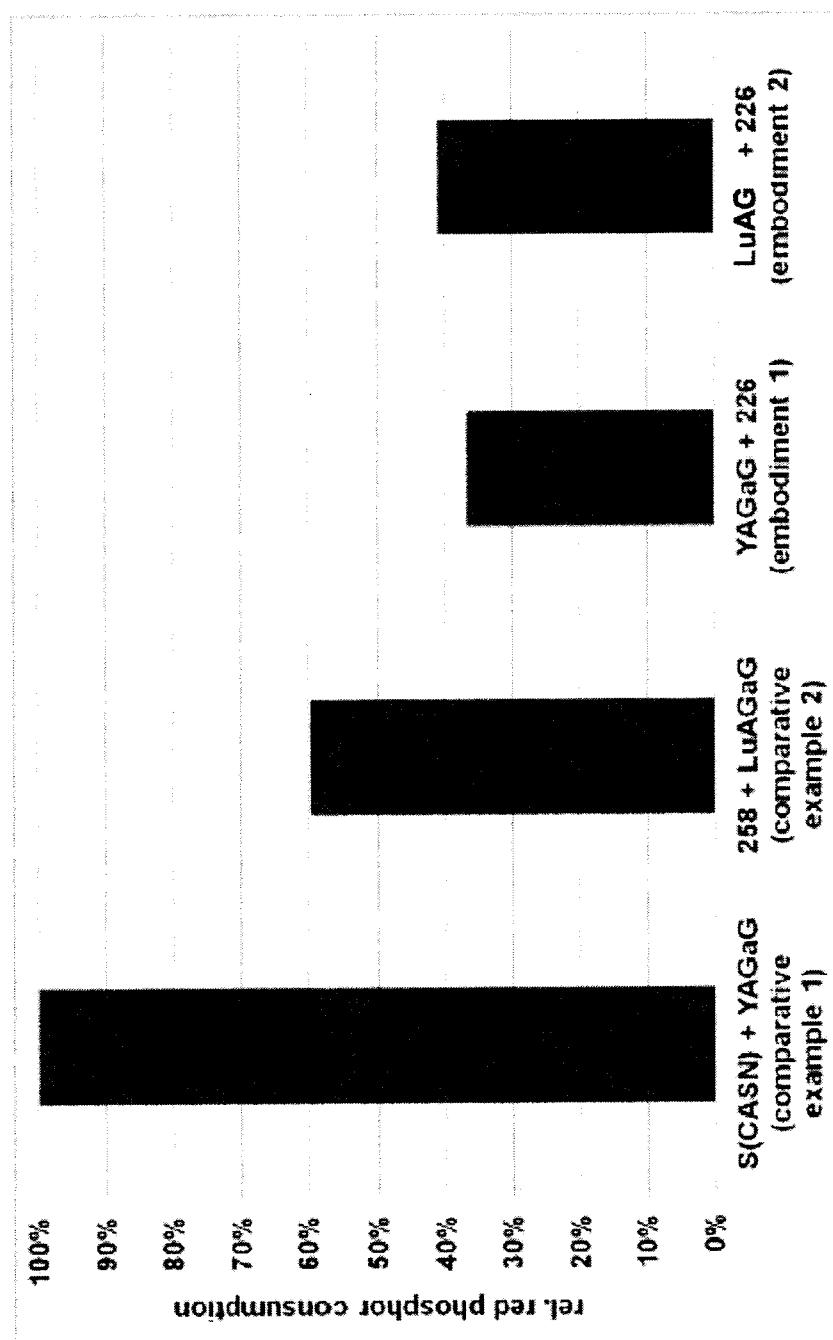


Fig. 51b

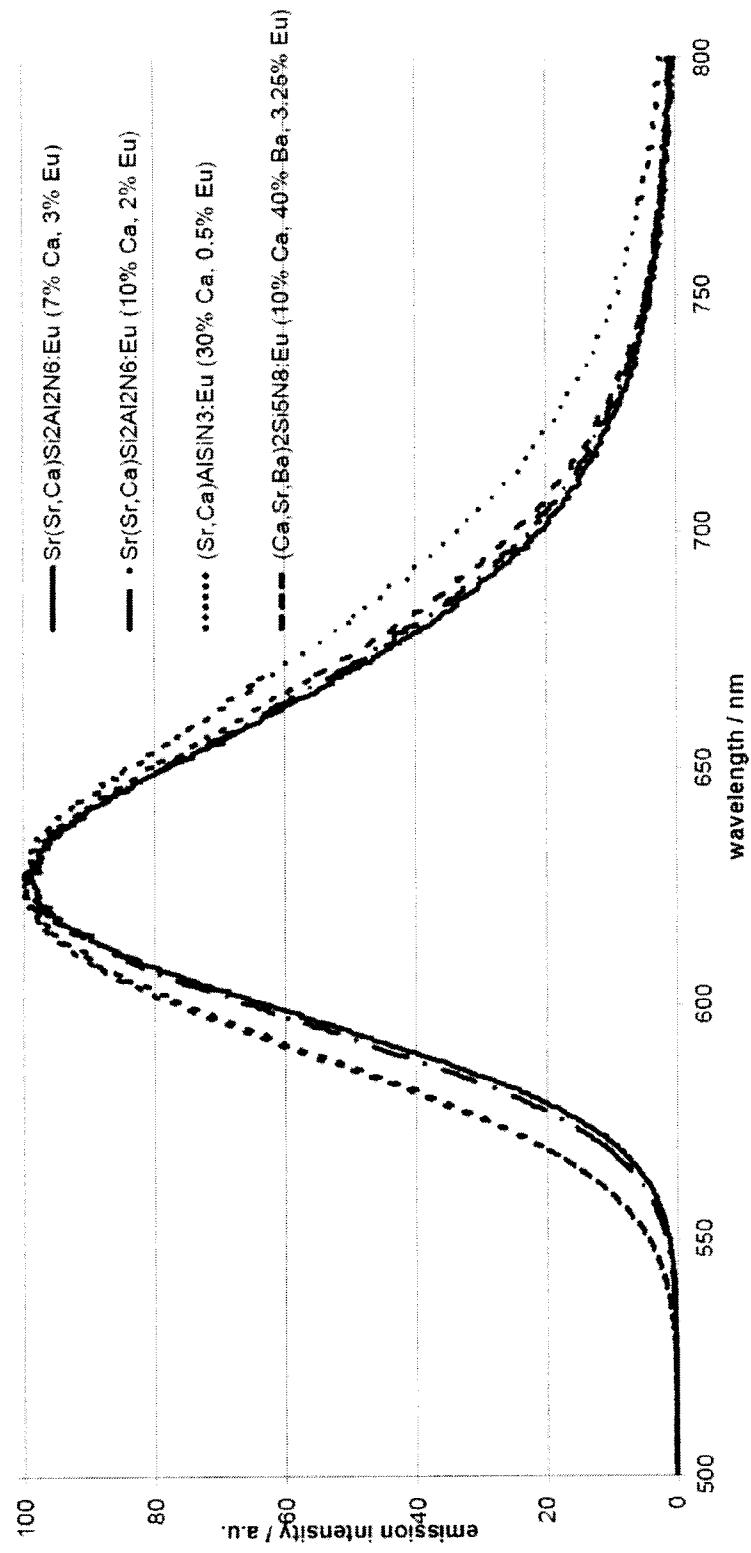


Fig. 51c

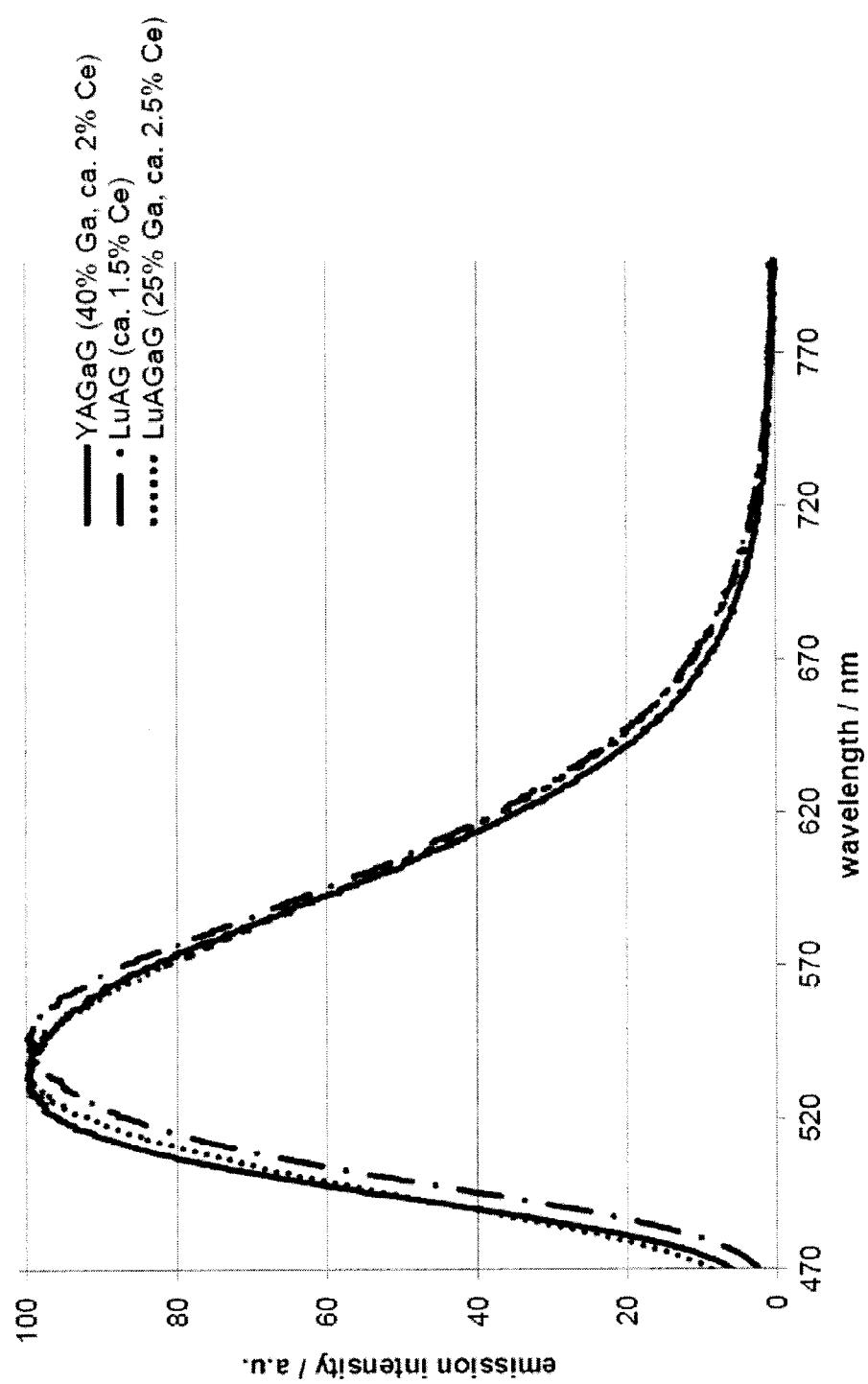


Fig. 51d

Fig. 51e

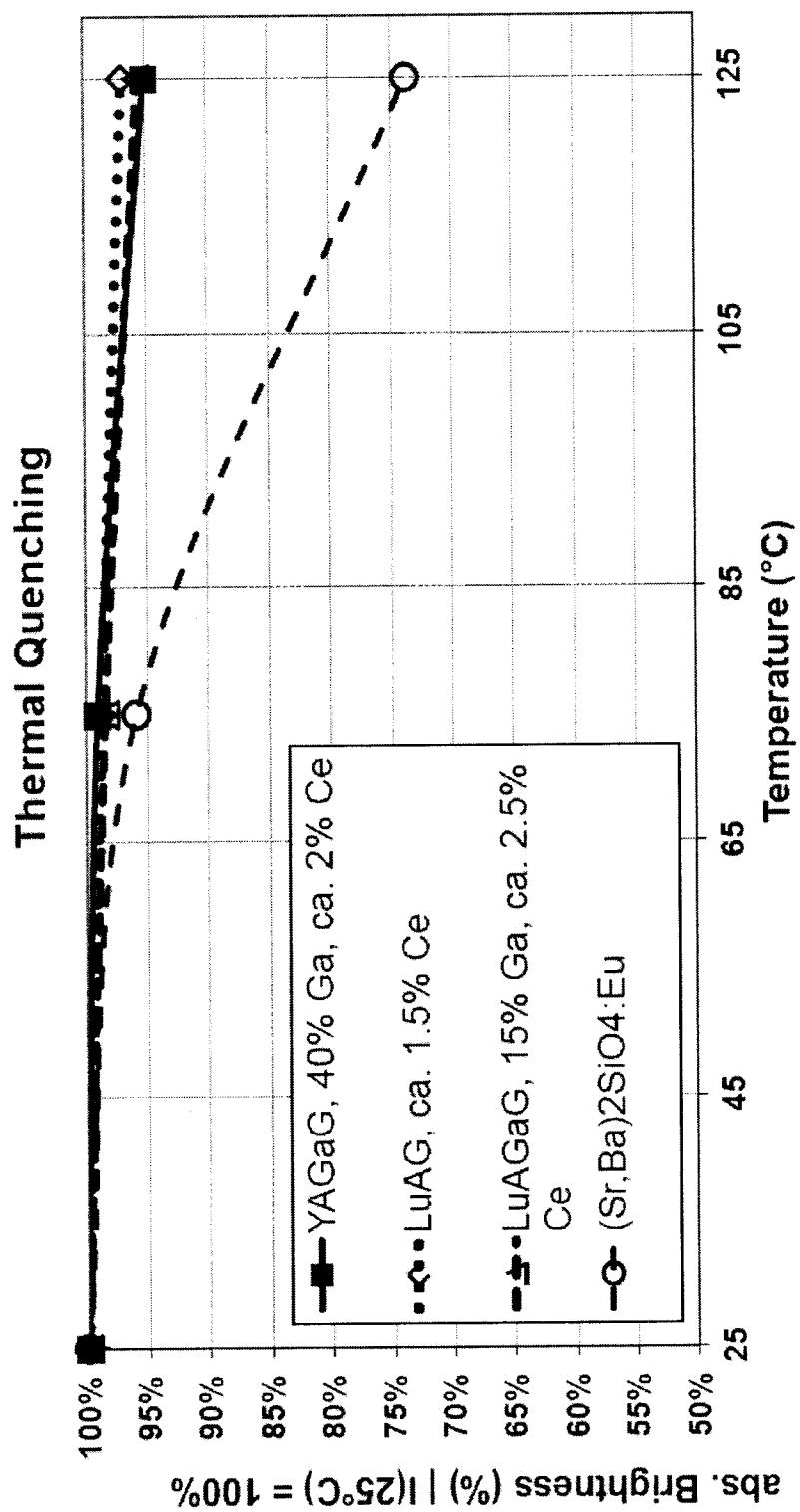
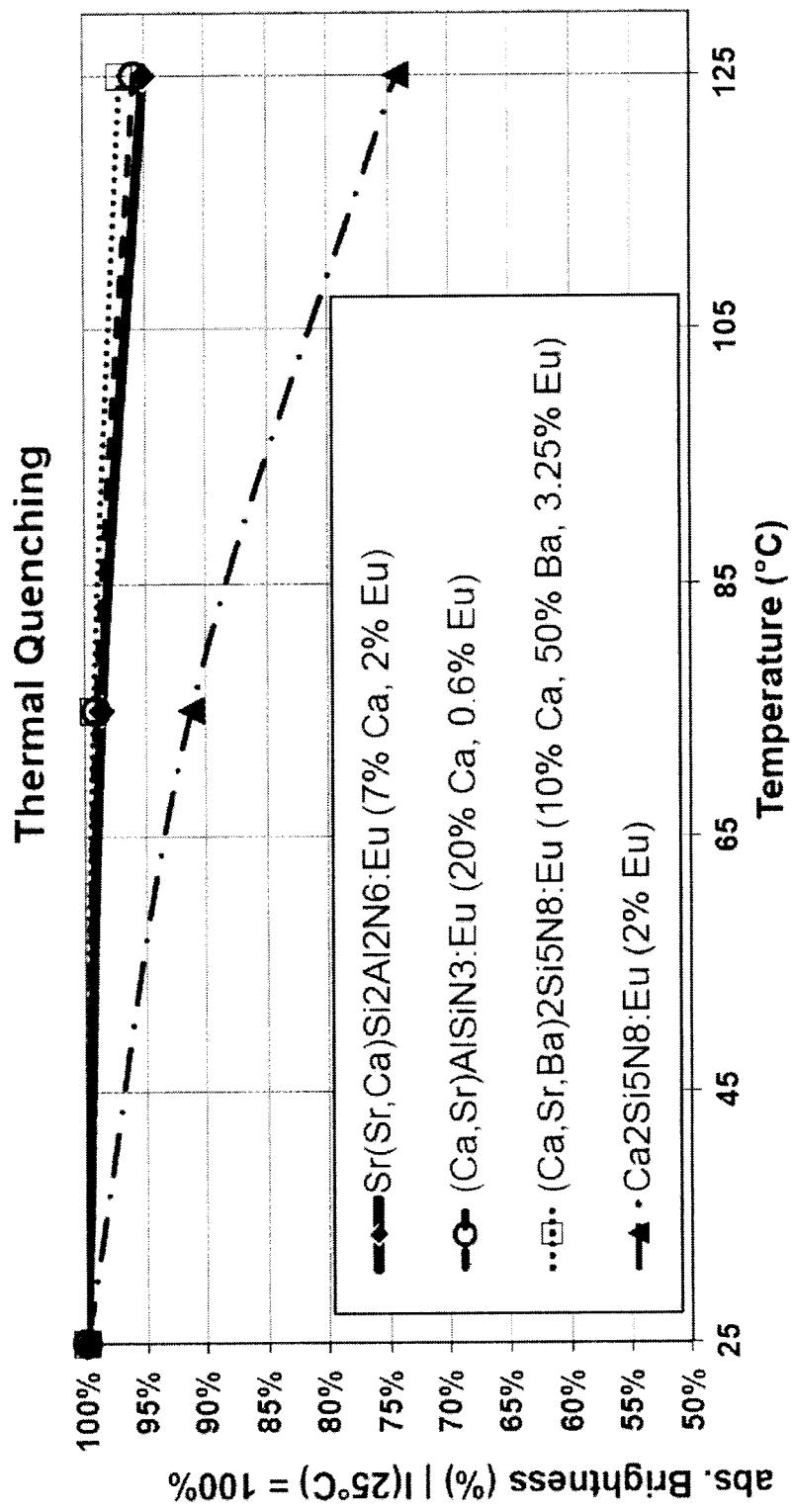


Fig. 51f



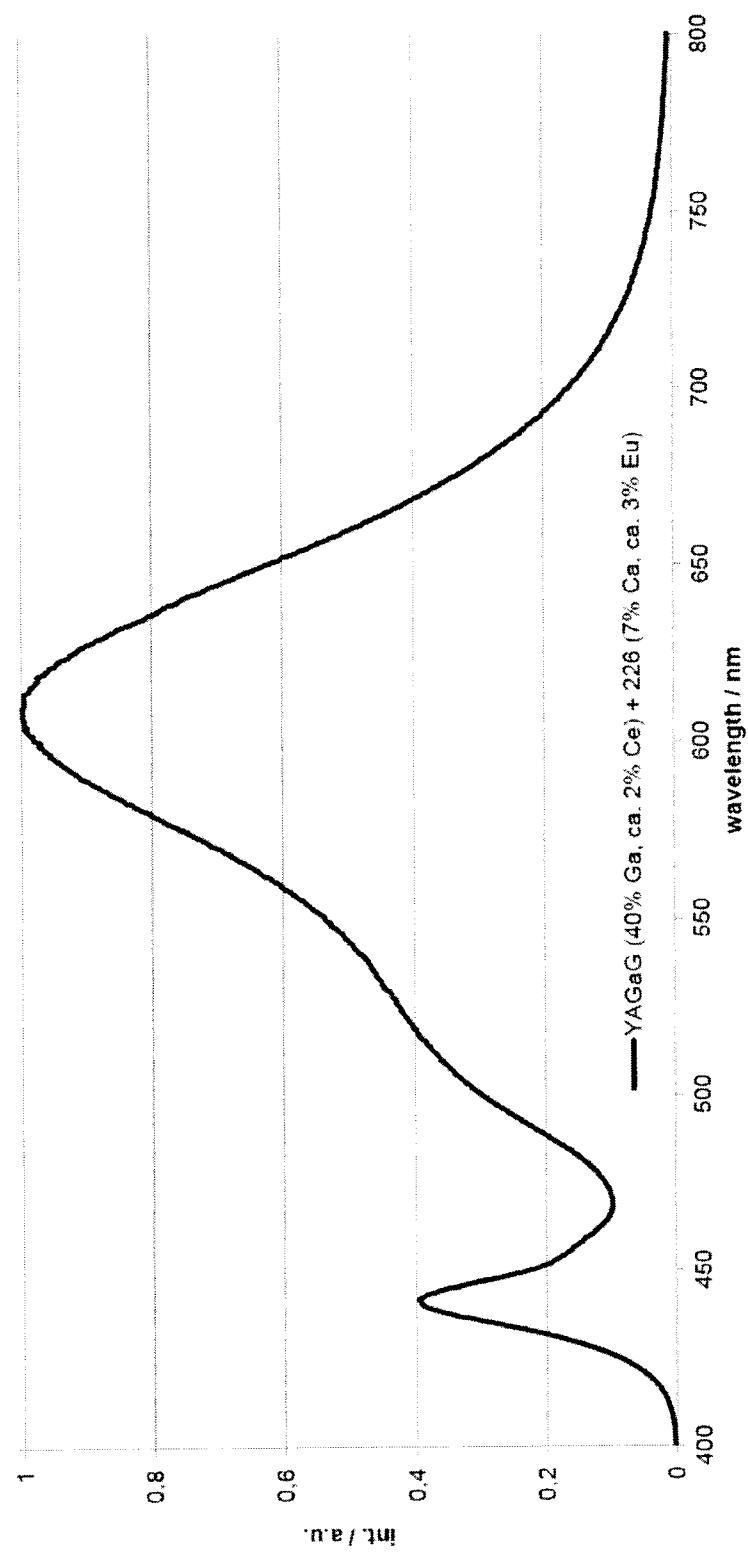


Fig. 51g

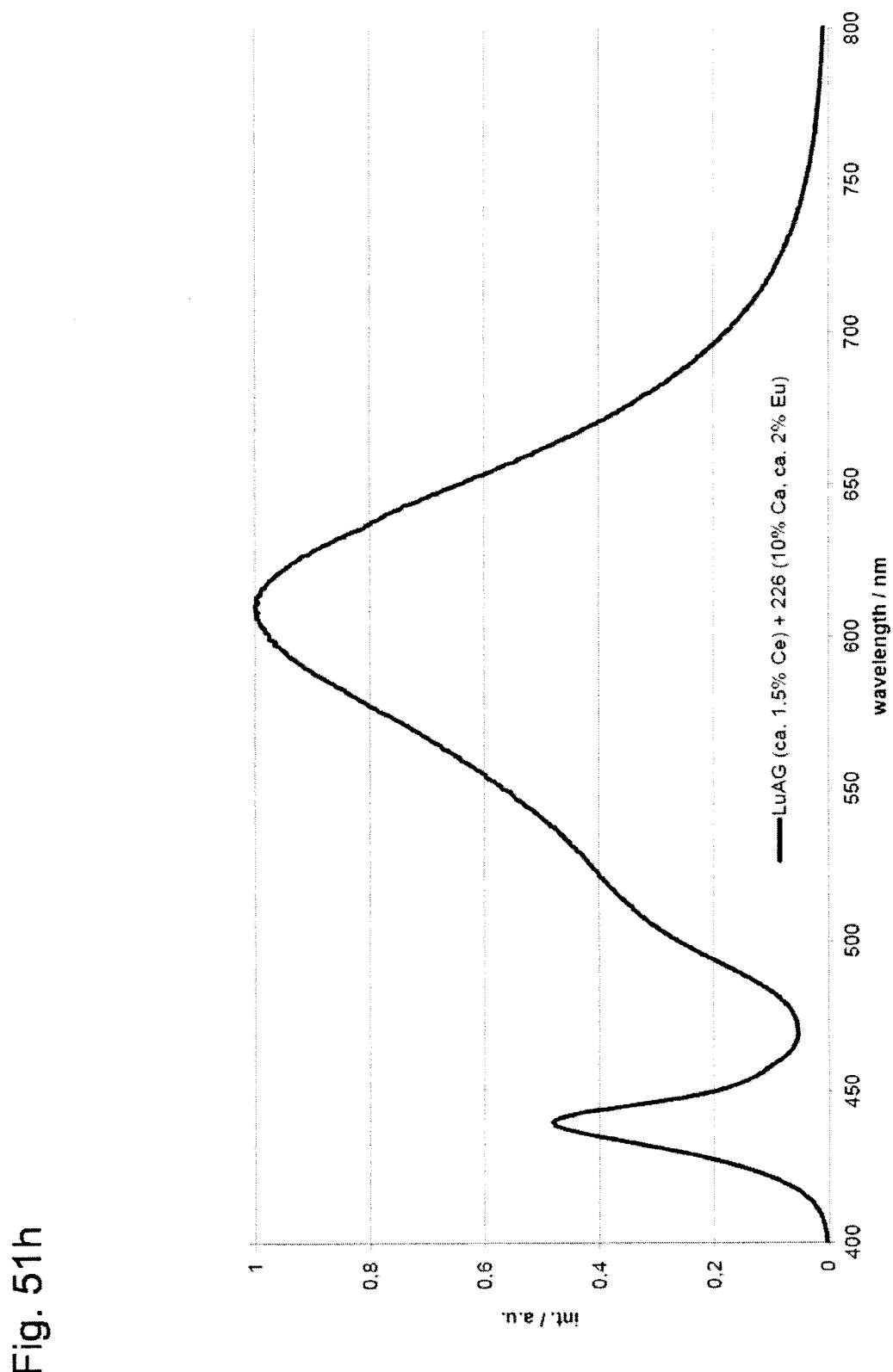


Fig. 52a

Examples	LED wave- length λ_{em} nm	Green		Red		Overall Ratio conc. g/r		Abs. green conc. %	Abs. red conc. %	CCT / K	CRI	Rel. conversion efficiency
		Conc.	Conc.	Conc.	Conc.	Conc.	Conc.					
Comparative example 1	446	(Lu,Ce) ₃ (Al,Ga) ₅ O ₁₂ (25% Ga, ca. 2.5% Ce)	(Sr _{0.4} Ca _{0.6})AlSiN ₃ :0.4% Eu	18,0%	2,78	13,2%	4,8%	2673	92	55	100%	
Inventive example 1	446	(Lu,Ce) ₃ (Al,Ga) ₅ O ₁₂ (25% Ga, ca. 2.5% Ce)	Sr(Sr _{0.84} Ca _{0.16})Al ₂ Si ₂ N ₆ :4.7% Eu	18,8%	12,60	17,4%	1,4%	2711	91	42	109%	
Inventive example 2	446	(Lu,Ce) ₃ (Al,Ga) ₅ O ₁₂ (25% Ga, ca. 2% Ce)	Sr(Sr _{0.84} Ca _{0.16})Al ₂ Si ₂ N ₆ :4.7% Eu	20,3%	15,00	19,0%	1,3%	2707	93	49	105%	
Inventive example 3	446	(Lu,Ce) ₃ (Al,Ga) ₅ O ₁₂ (25% Ga, ca. 2% Ce)	Sr(Sr _{0.8} Ca _{0.2})Al ₂ Si ₂ N ₆ :3.7% Eu	20,5%	12,40	19,0%	1,5%	2670	91	40	107%	
Inventive example 4	446	(Lu,Ce) ₃ (Al,Ga) ₅ O ₁₂ (25% Ga, ca. 1.8% Ce)	Sr(Sr _{0.8} Ca _{0.2})Al ₂ Si ₂ N ₆ :3.7% Eu	20,5%	13,00	19,0%	1,5%	2726	92	44	109%	

Fig. 52b

Example	LED wavelength $\lambda_{\text{dom}} / \text{nm}$	Green	Red	Overall Ratio conc. g/r	Abs. green conc.	Abs. red conc.	CCT / K	CRI	R9	Conversion efficiency	
Comparative example 1	446	(Lu,Ce)3(Al,Ga)5O12 (25% Ga, ca. 2.5% Ce)	(Sr0.4Ca0.6)AlSiN3:0.4% Eu	18.0%	2.78	13.2%	4,8%	2673	92	55	100%
Inventive example 1	446	(Lu,Ce)3(Al,Ga)5O12 (25% Ga, ca. 2.5% Ce)	Sr(Sr0.8Ca0.2)Al2Si2N6:4.7% Eu	18.8%	12.60	17.4%	1,4%	2711	91	42	109%
Comparative example 1	446	(Lu,Ce)3(Al,Ga)5O12 (25% Ga, ca. 2.5% Ce)	(Sr0.4Ca0.6)AlSiN3:0.4% Eu	13.2%	3.87	10.5%	2,7%	3974	90	61	100%
Inventive example 1	446	(Lu,Ce)3(Al,Ga)5O12 (25% Ga, ca. 2.5% Ce)	Sr(Sr0.8Ca0.2)Al2Si2N6:4.7% Eu	13.5%	16.00	12.7%	0,8%	3986	90	50	106%

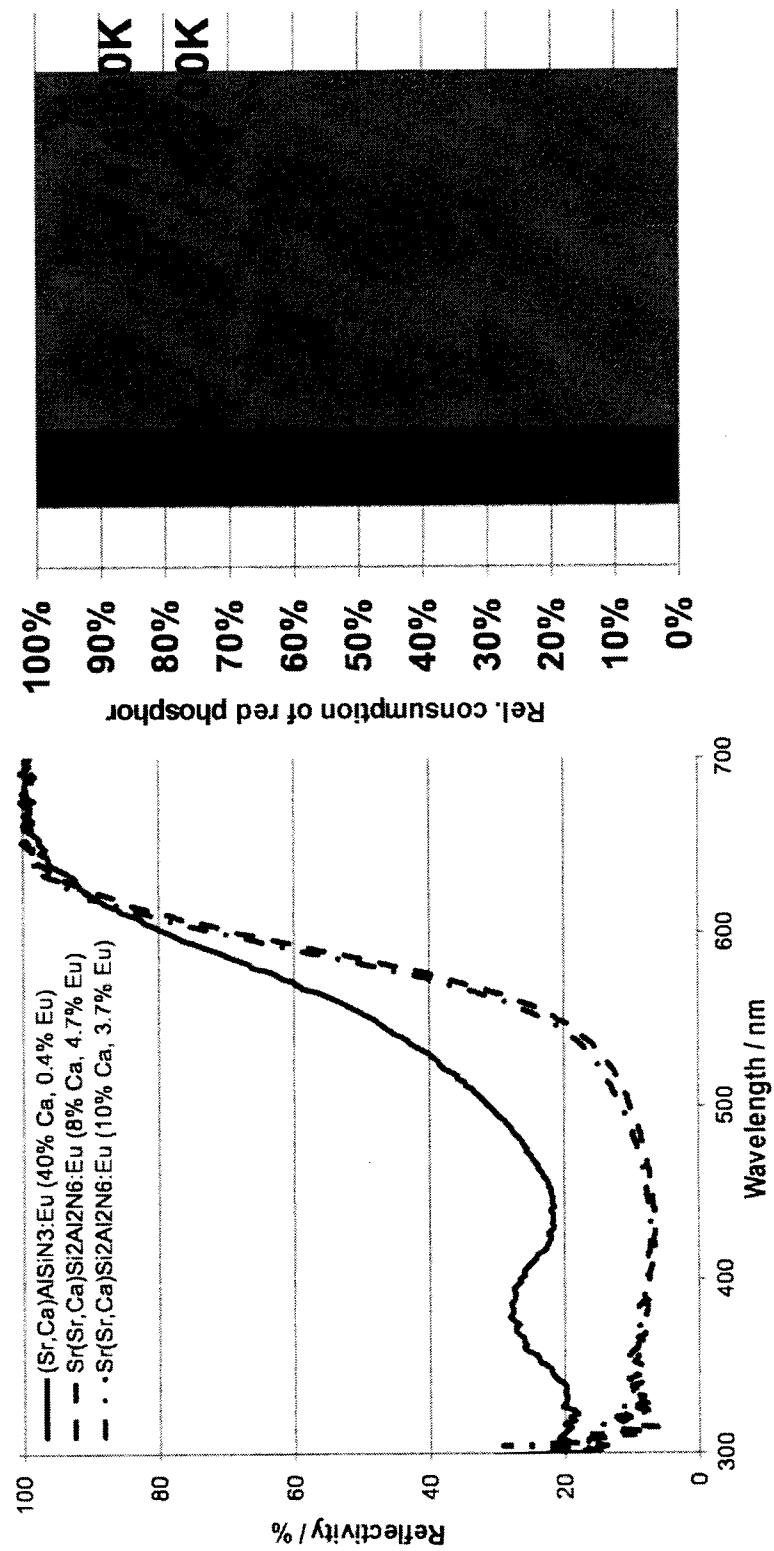
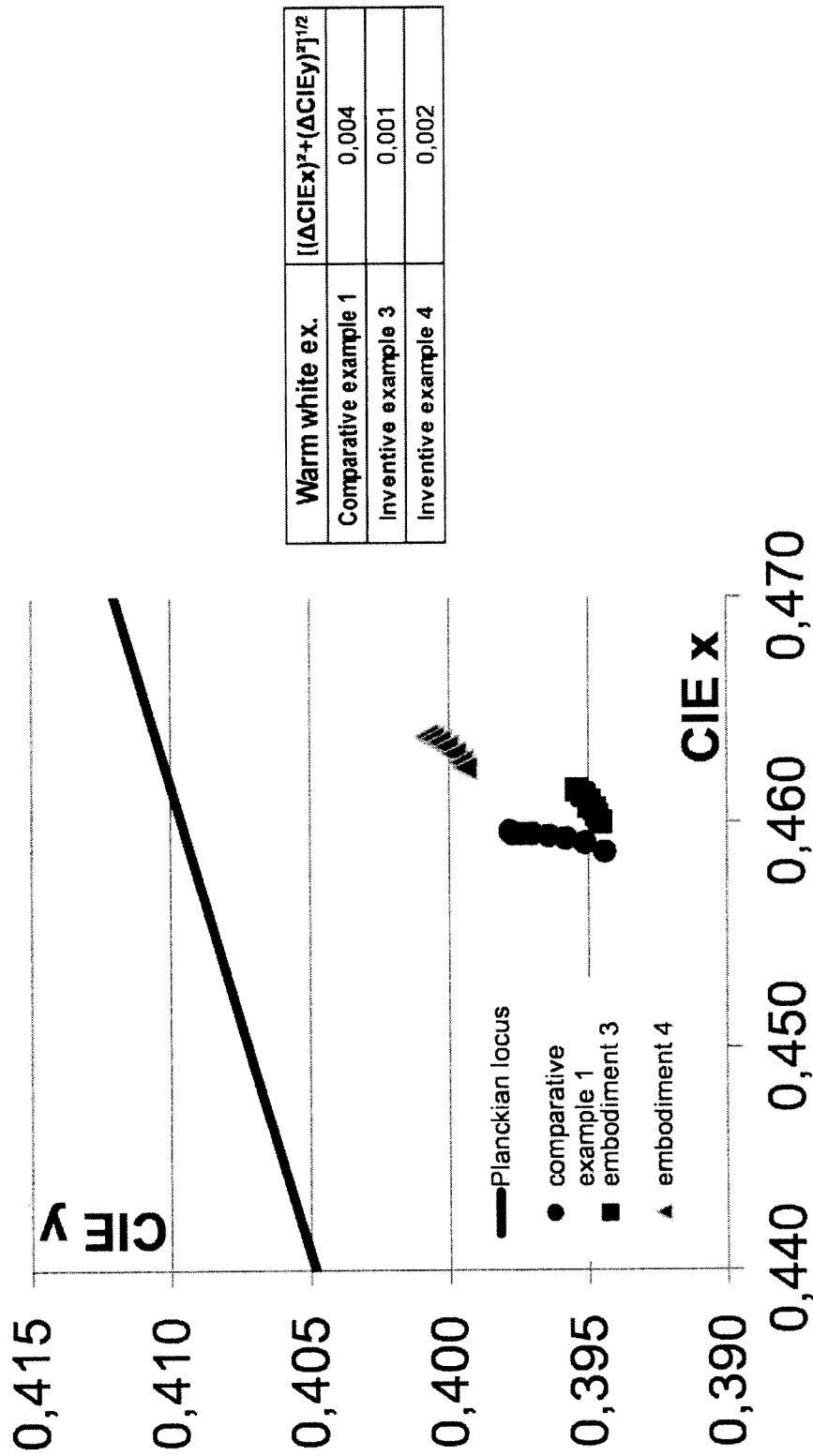


Fig. 52c

Fig. 52d



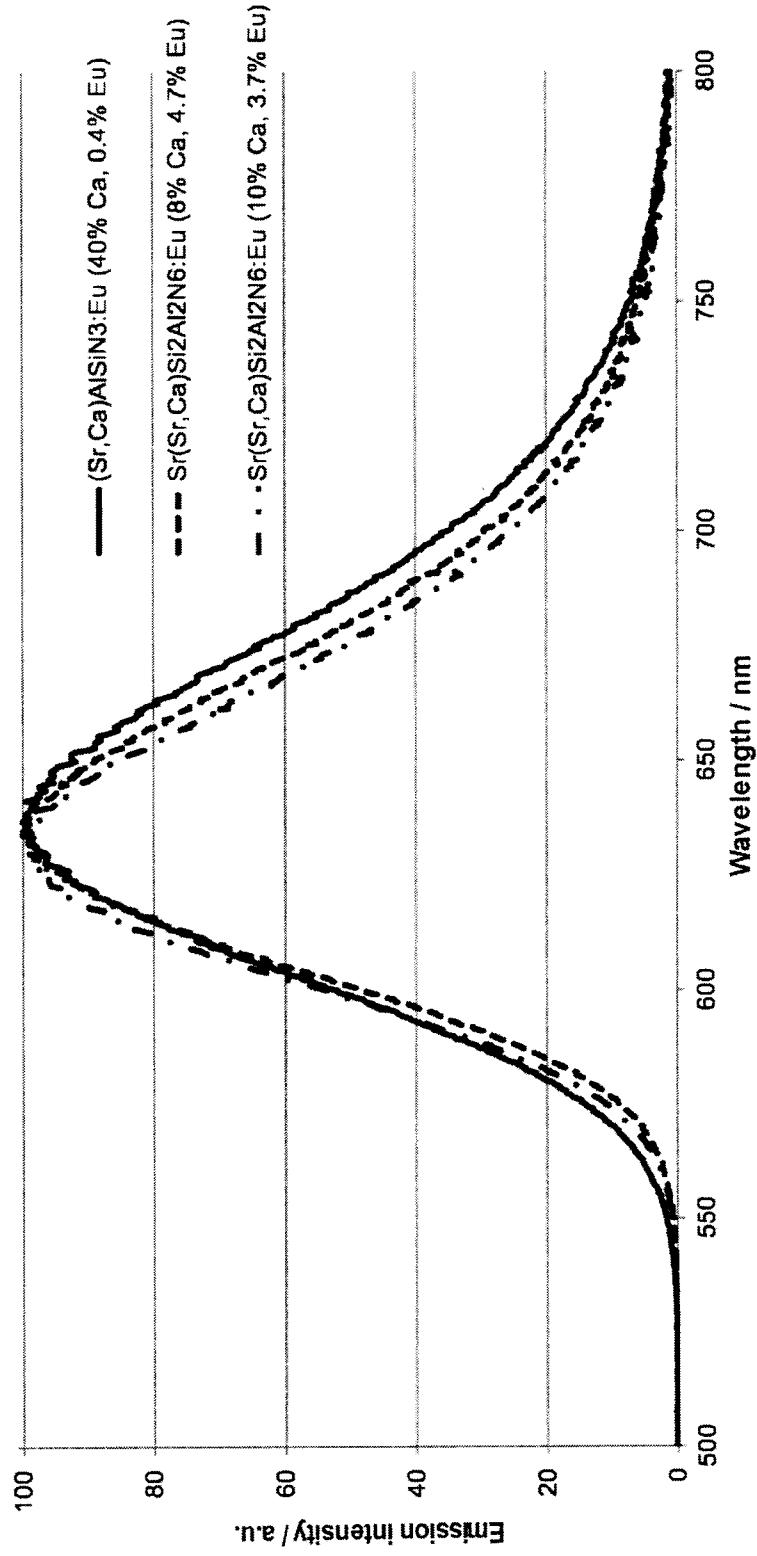


Fig. 52e

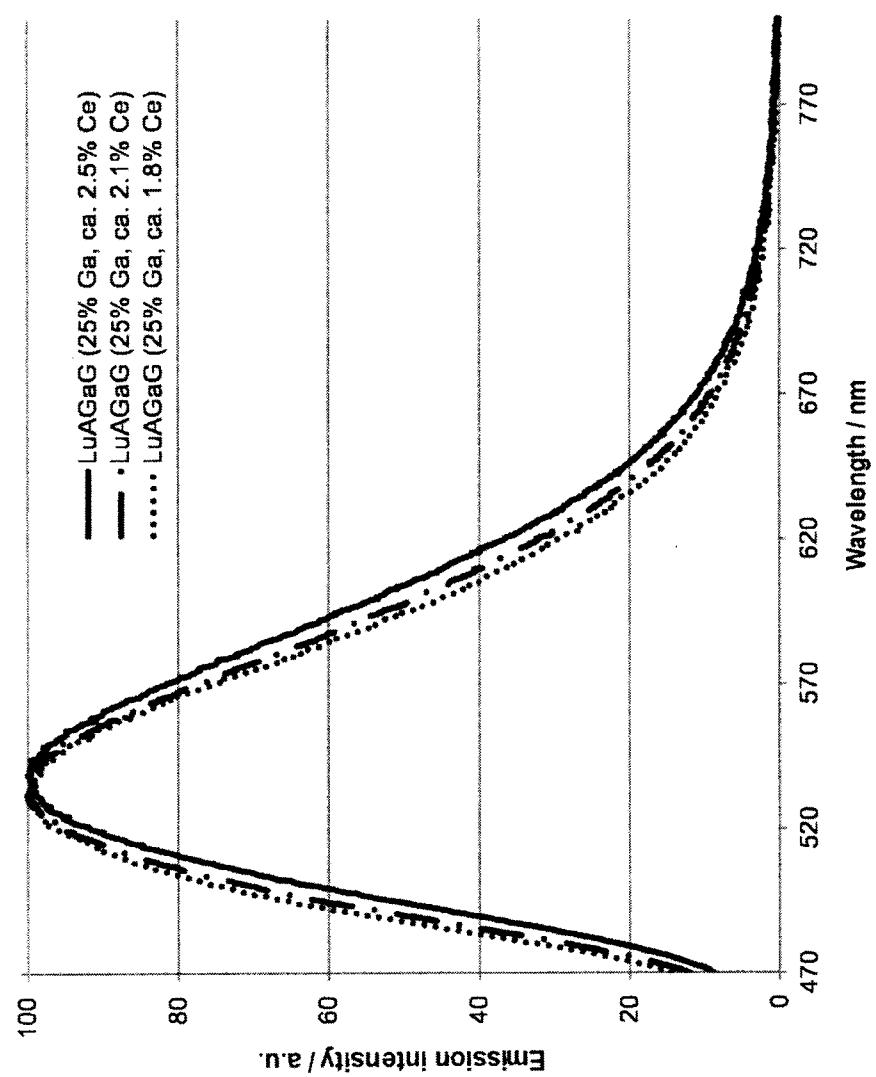


Fig. 52f

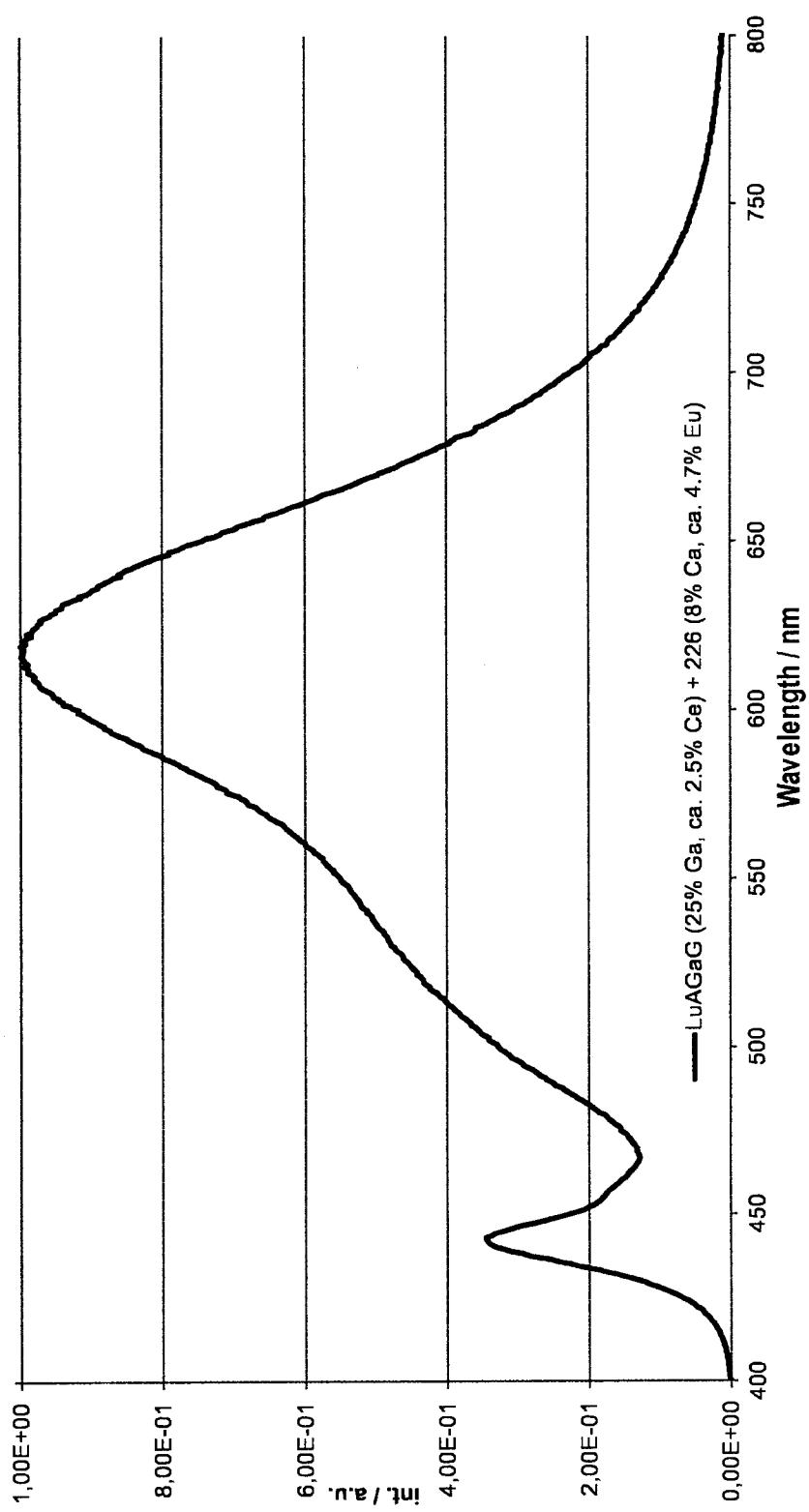


Fig. 52g

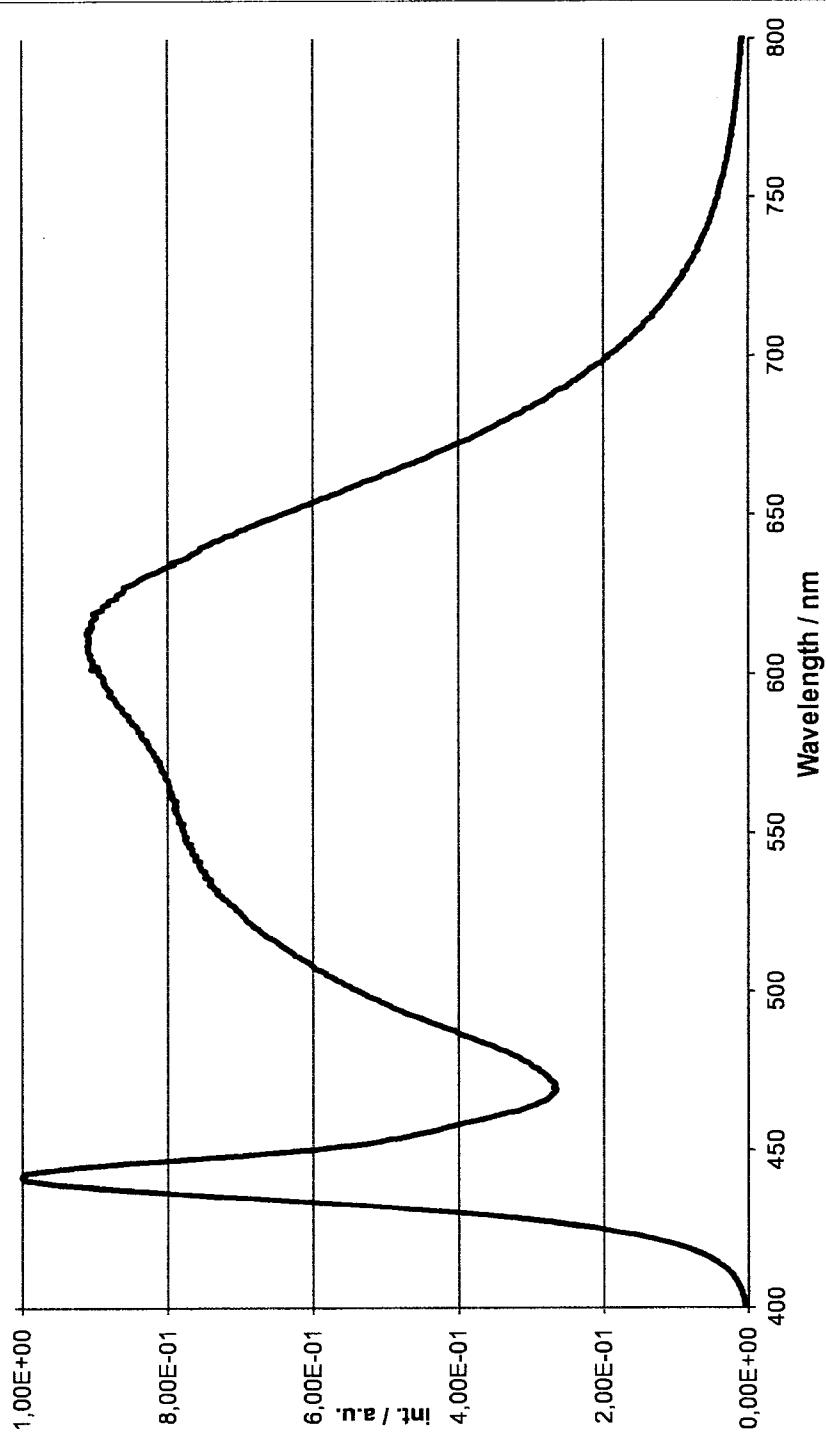


Fig. 52h

Fig. 53a

Example	$\lambda_{\text{emission}}$ LED at 360 mA	Composition of yellow-green phosphor	Composition of orange-red phosphor	Composition of red phosphor	Total phosphor conc. (% by wt.)	Yellow-green/red or yellow-green/ red/red ratio
Comp. ex. 1	447	$\text{Lu}_3(\text{Al}_{0.75}\text{Ga}_{0.25})_5\text{O}_1$ 2% Ce (2% Eu)	$\text{CaAlSi}(\text{N},\text{O})_3\text{:Eu}$ (0.4% Eu)		21.5%	2.4:1
Comp. ex. 2	447	$\text{Lu}_3(\text{Al}_{0.75}\text{Ga}_{0.25})_5\text{O}_1$ 2% Ce (2.0 % Ce)	$(\text{Ca}_{0.025}\text{Sr}_{0.475}\text{Ba}_{0.5})$ $2\text{Si}_5\text{N}_6\text{:Eu}$ (2.5% Eu)	$\text{CaAlSi}(\text{N},\text{O})_3\text{:}$ Eu (0.4% Eu)	19.0%	3.6:0.2:0.8
Inventive ex. 1	447	$\text{Lu}_3(\text{Al}_{0.75}\text{Ga}_{0.25})_5\text{O}_1$ 2% Ce (2% Eu)	$(\text{Sr}_{0.5}\text{Ba}_{0.5})_2$ $\text{Si}_5\text{N}_6\text{:Eu}$ (1% Eu)	$\text{Sr}(\text{Sr}_{0.8}\text{Ca}_{0.2})$ $\text{Si}_2\text{Al}_2\text{N}_6\text{:Eu}$ (4.5% Eu)	20.6%	13.7:0.14:0.86
Inventive ex. 2	447	$\text{Lu}_3(\text{Al}_{0.75}\text{Ga}_{0.25})_5\text{O}_1$ 2% Ce (2 % Ce)	$\text{Sr}(\text{Sr}_{0.88}\text{Ca}_{0.14})\text{Si}_2\text{A}$ $\text{I}_2\text{N}_6\text{:Eu}$ (0.8% Eu)	$\text{Sr}(\text{Sr}_{0.80}\text{Ca}_{0.2})$ $\text{Si}_2\text{Al}_2\text{N}_6\text{:Eu}$ (4.5% Eu)	20.6%	13.6:0.11:0.89

Example	CIE-x	CIE-y	CRI	Rel. Φ_v (filled potting)/ Φ _e (clear potting)	Rel. LER (luminous efficiency)	Rel. Φ_e (filled potting)/ Φ _e (clear potting)
Comp. ex. 1	0,460	0,411	97	100%	100%	100%
Comp. ex. 2	0,459	0,412	94	107,5%	105,9%	101,5%
Inventive ex. 1	0,460	0,413	92	120,9%	115,7%	104,5%
Inventive ex. 2	0,460	0,409	92	120,2%	114,7%	104,8%

Fig. 53b

Fig. 53c

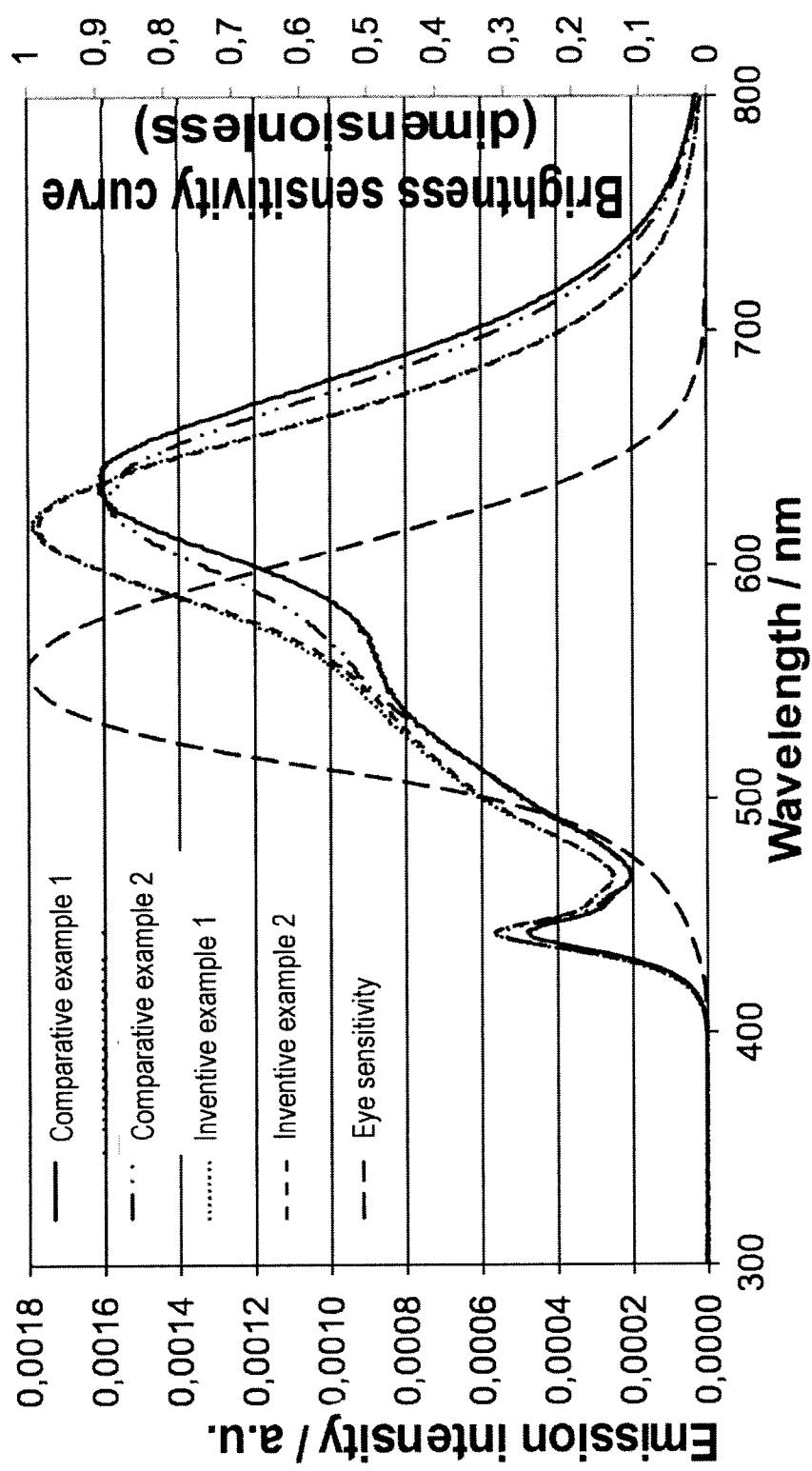
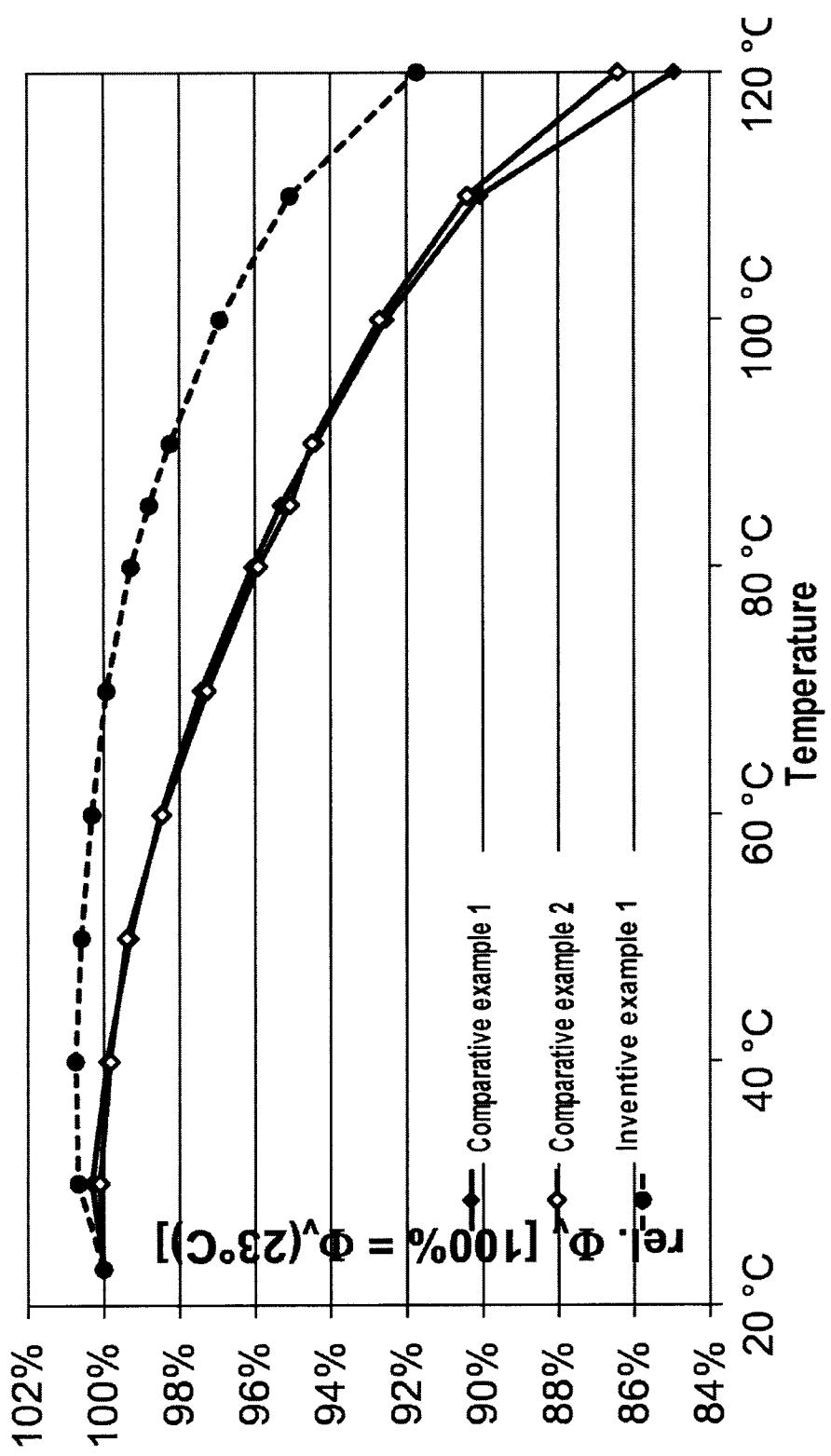


Fig. 53d



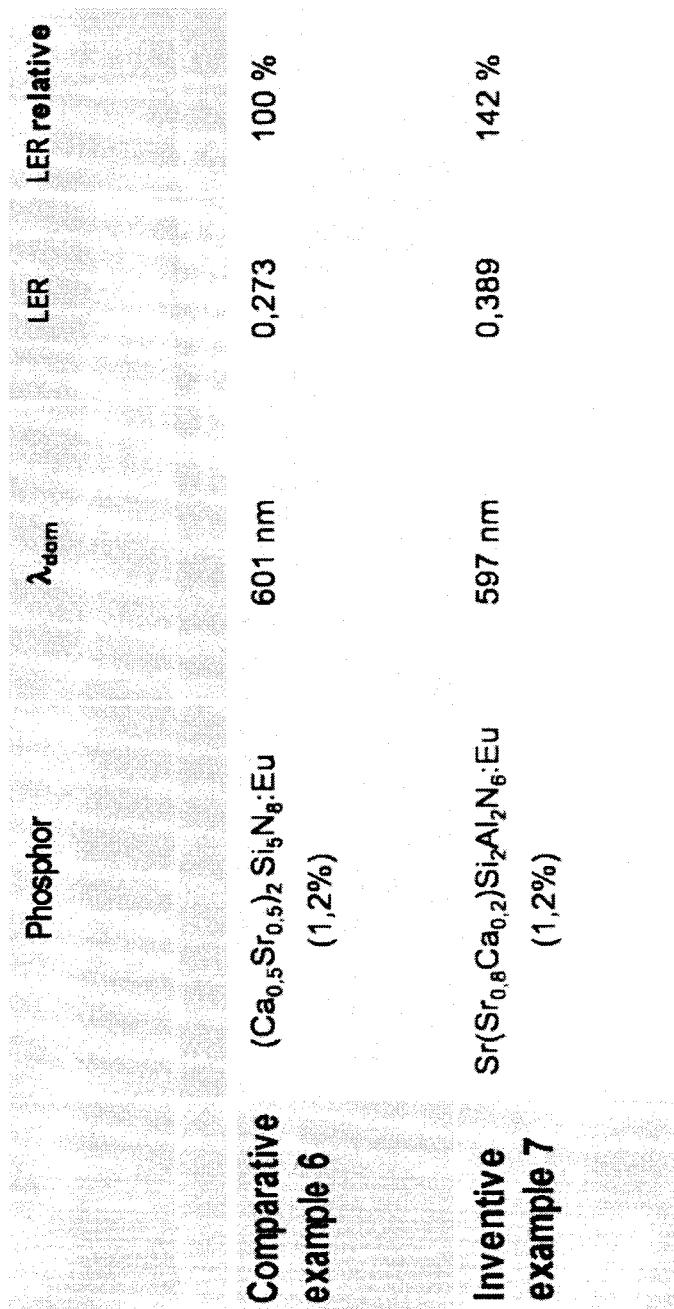


Fig. 54

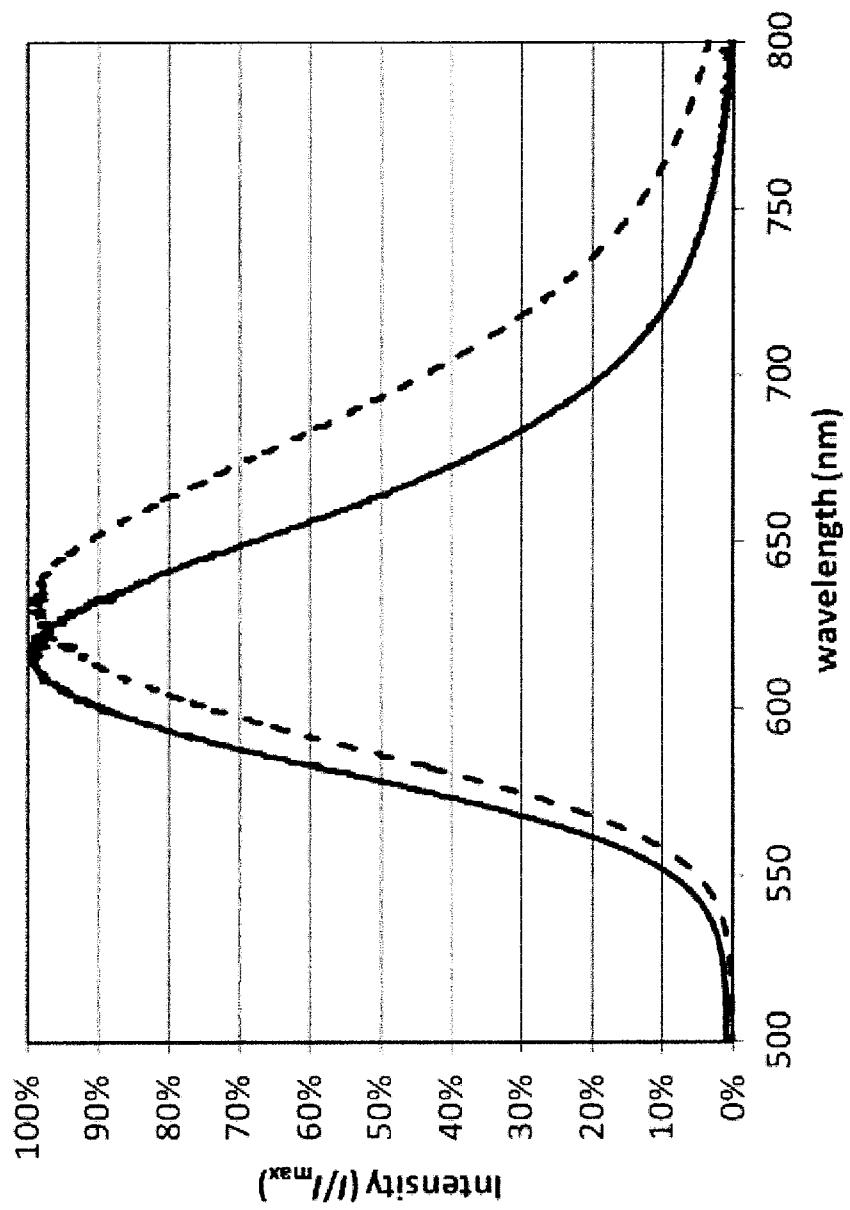
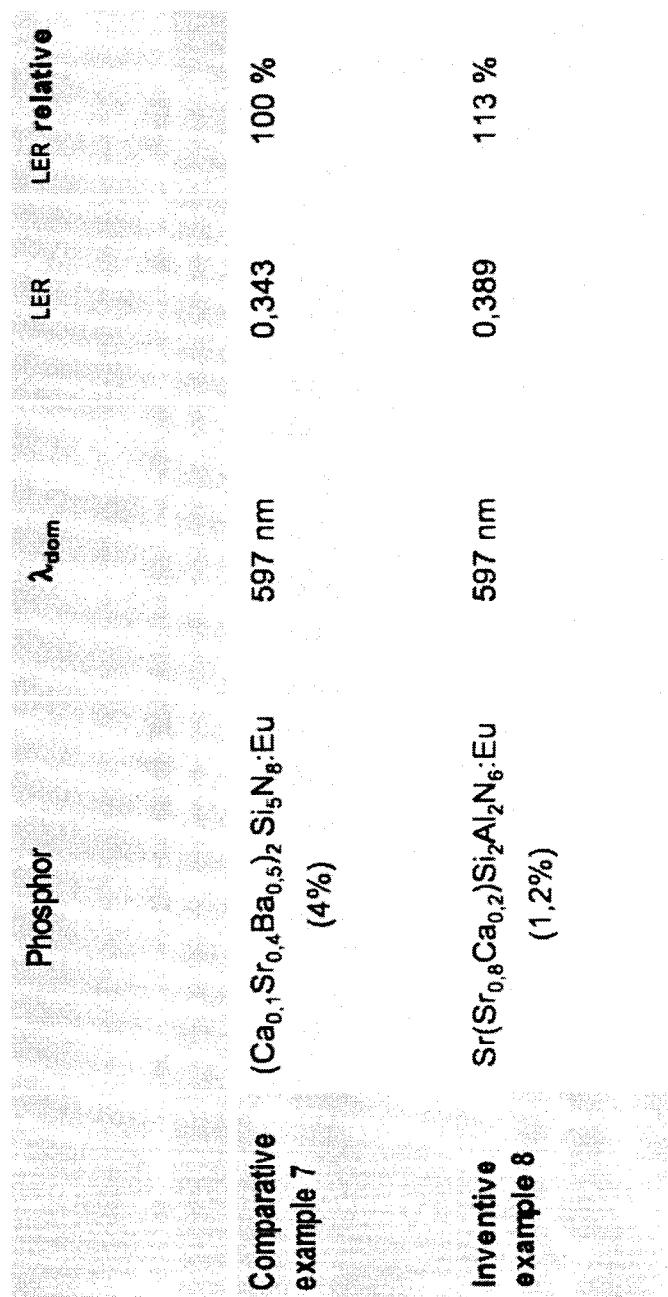


Fig. 55

Fig. 56



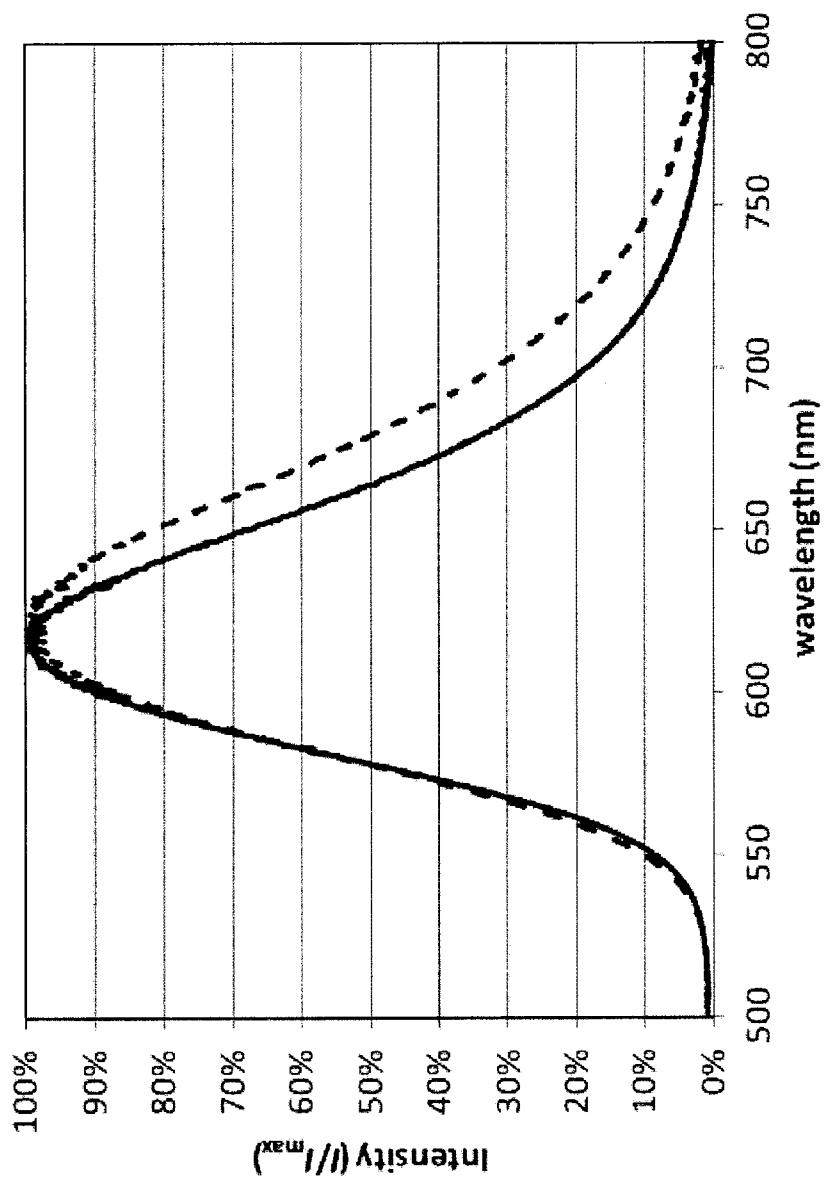
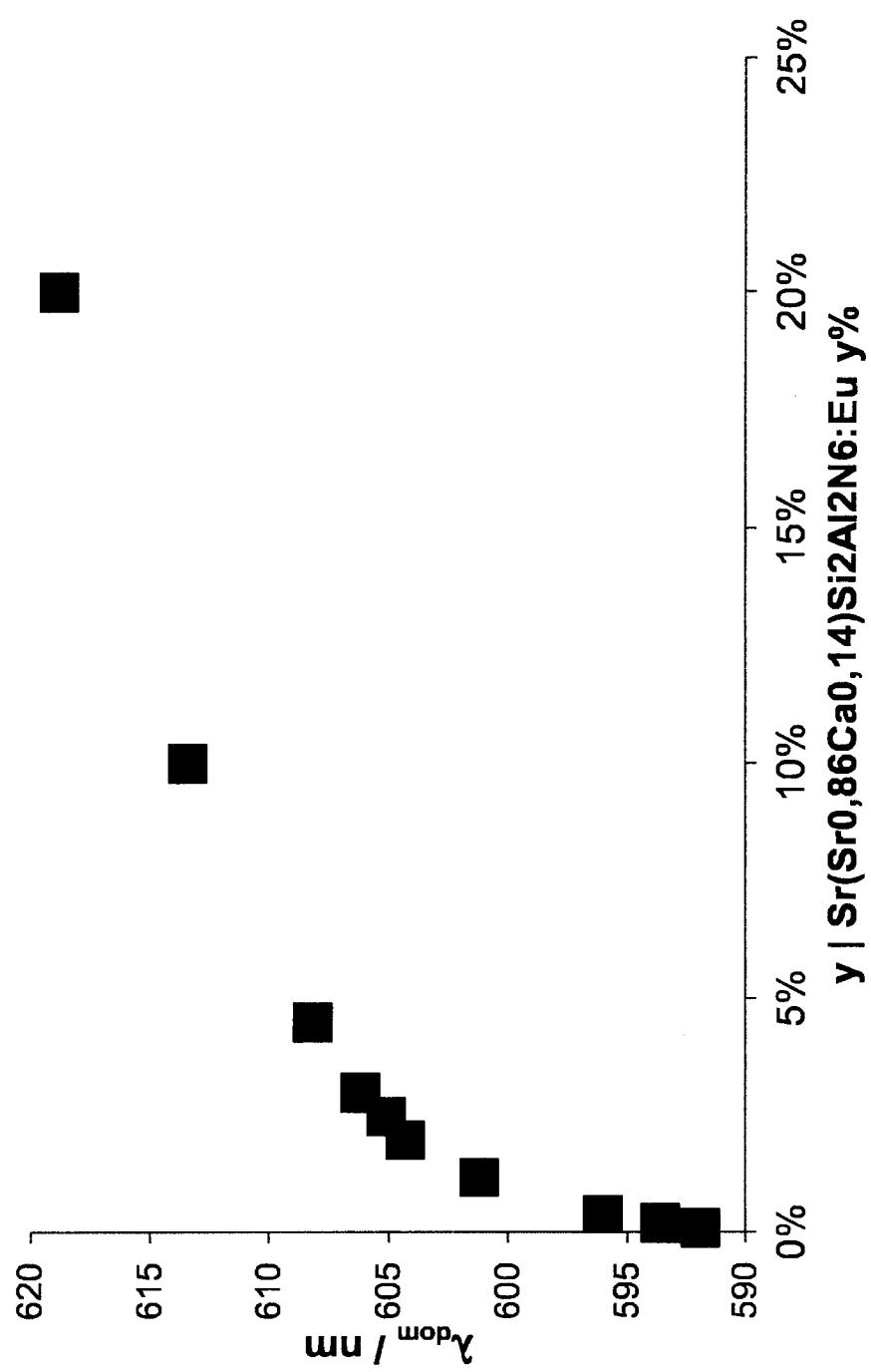


Fig. 57

Fig. 58a



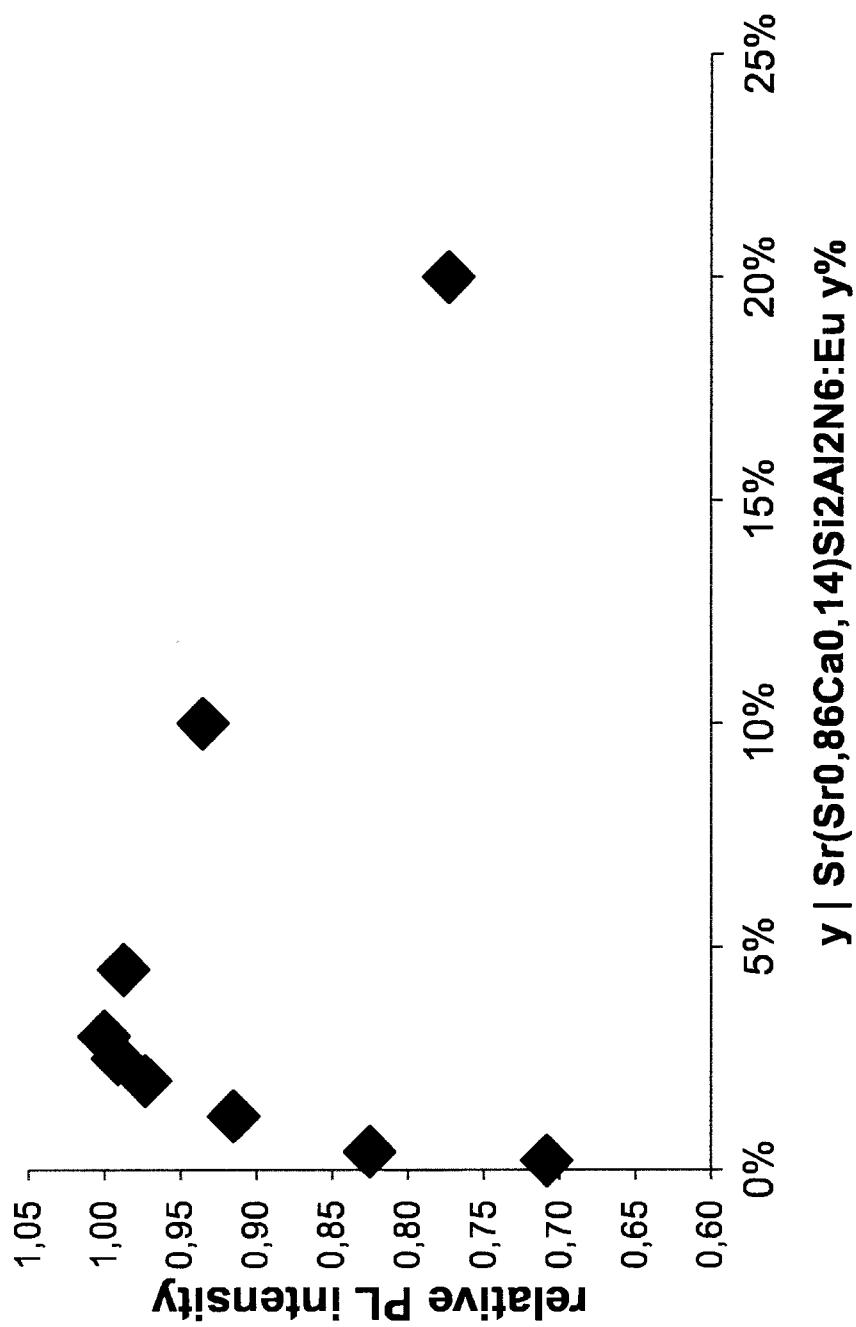


Fig. 58b

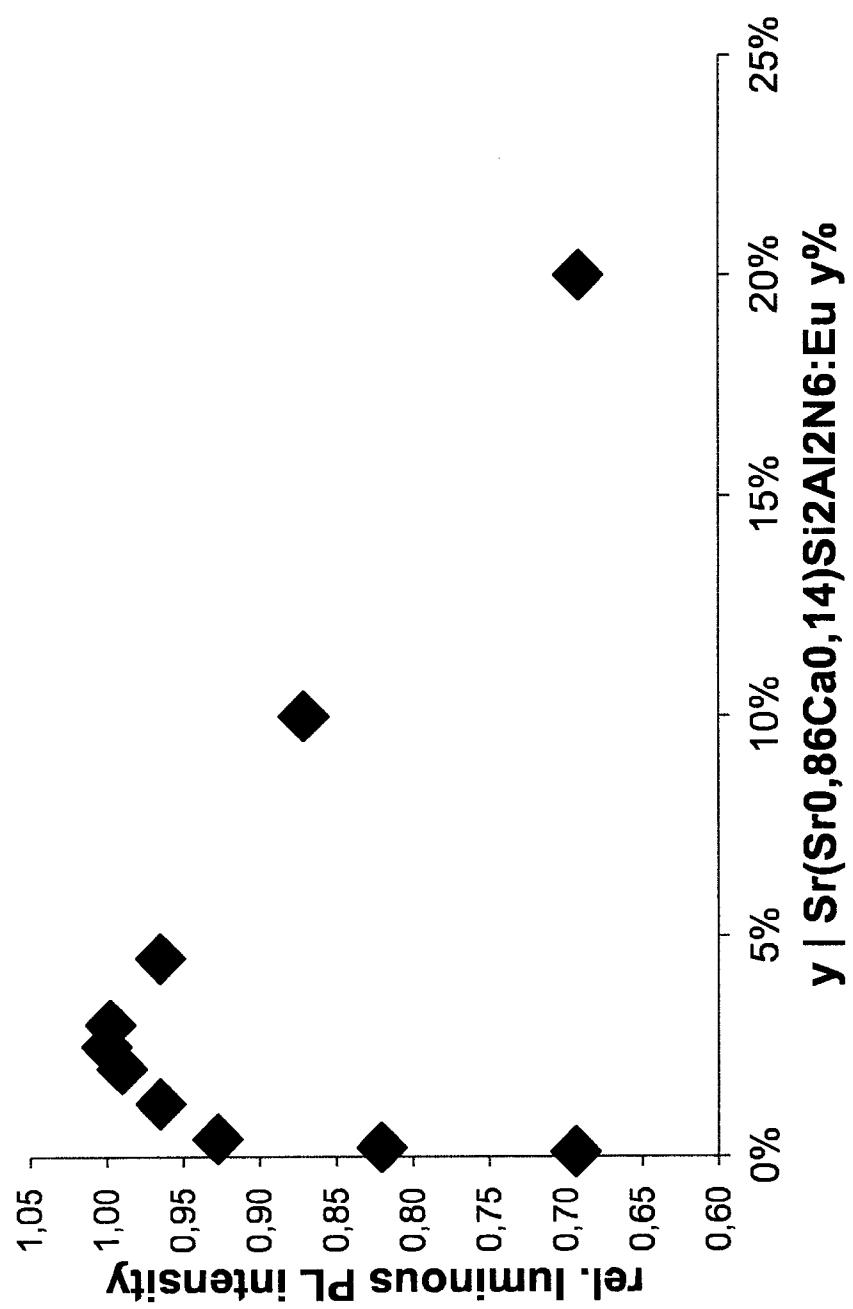


Fig. 58c

Fig. 59a

Nominal composition	R(450-470)	CIE-x:	CIE-y:	LER	l _{dom} [nm]	FWHM [nm]
$Sr_{(1-x)}(Sr_{0.8}Ca_{0.2})_{(1-y)}Cu_{(x+y)}Si_2Al_2N_6 : 2\% Eu (x+y=0.2)$	15,6%	0,646	0,352	0,260	604,8	90,0
$Sr_{(1-x)}(Sr_{0.8}Ca_{0.2})_{(1-y)}7n_{(x+y)}Si_2Al_2N_6 : 2\% Eu (x+y=0.2)$	14,8%	0,648	0,350	0,275	605,4	83,2
$Sr_{(1-x)}(Sr_{0.8}Ca_{0.2})_{(1-y)}(LaLi)_{(x+y)}Si_2Al_2N_6 : 2\% Eu (x+y=0.2)$	11,4%	0,640	0,353	0,183	604,5	146,0
$Sr_{(1-x)}(Sr_{0.86}Ca_{0.14})_{(1-y)}(LaLi)_{(x+y)}Si_2Al_2N_6 : 3\% Eu (x+y=0.04)$	10,6%	0,648	0,349	0,215	605,6	103,3
$Sr_{(1-x)}(Sr_{0.86}Ca_{0.14})_{(1-y)}La_{(x+y)}Si_2-(x+y)Al_2-(x+y)N_6 : 3\% Eu (x+y=0.2)$	10,0%	0,464	0,352	0,234	604,8	100,3

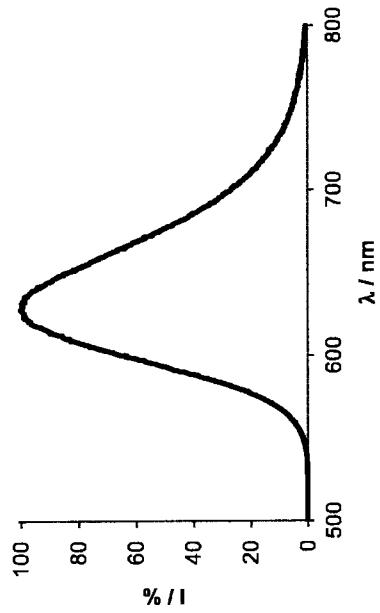
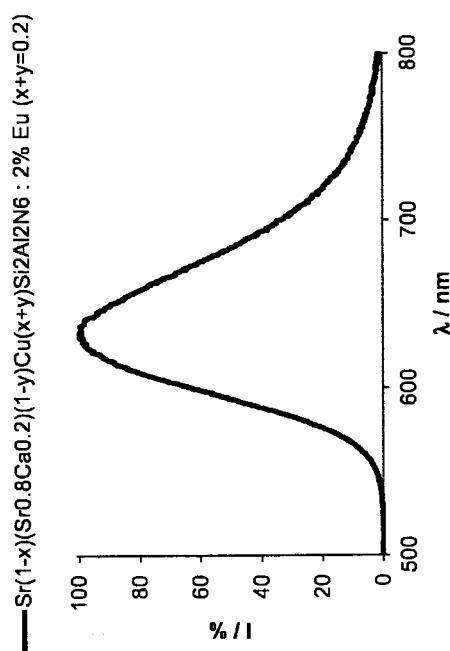


Fig. 59d $\text{Sr}(1-x)(\text{Sr}0.8\text{Ca}0.2)(1-y)(\text{LaLi})(x+y)\text{Si}2\text{Al}2\text{N}6 : 2\% \text{Eu} (x+y=0.2)$

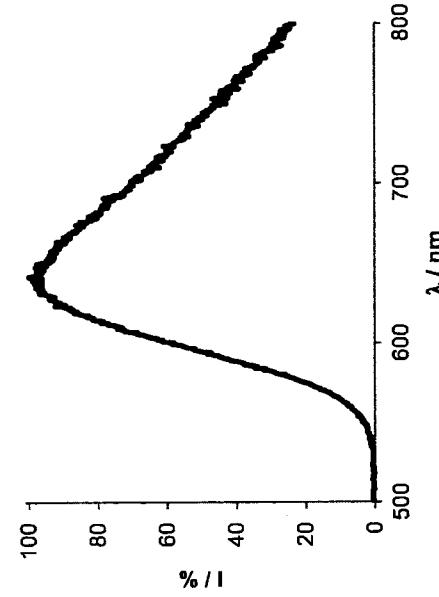


Fig. 59e $\text{Sr}(1-x)(\text{Sr}0.86\text{Ca}0.14)(1-y)(\text{LaLi})(x+y)\text{Si}2\text{Al}2\text{N}6 : 3\% \text{Eu} (x+y=0.04)$

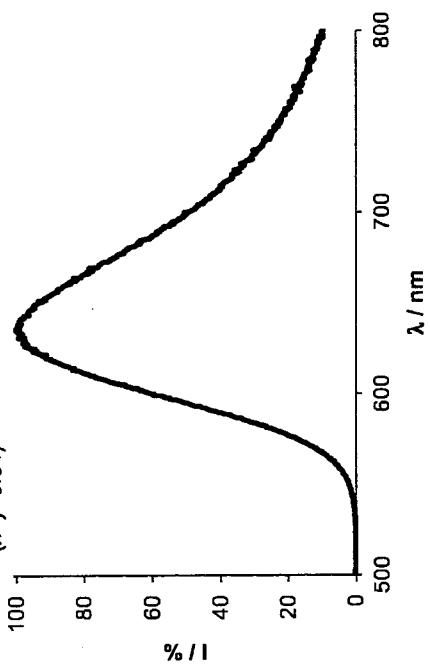


Fig. 59d

Fig. 59e

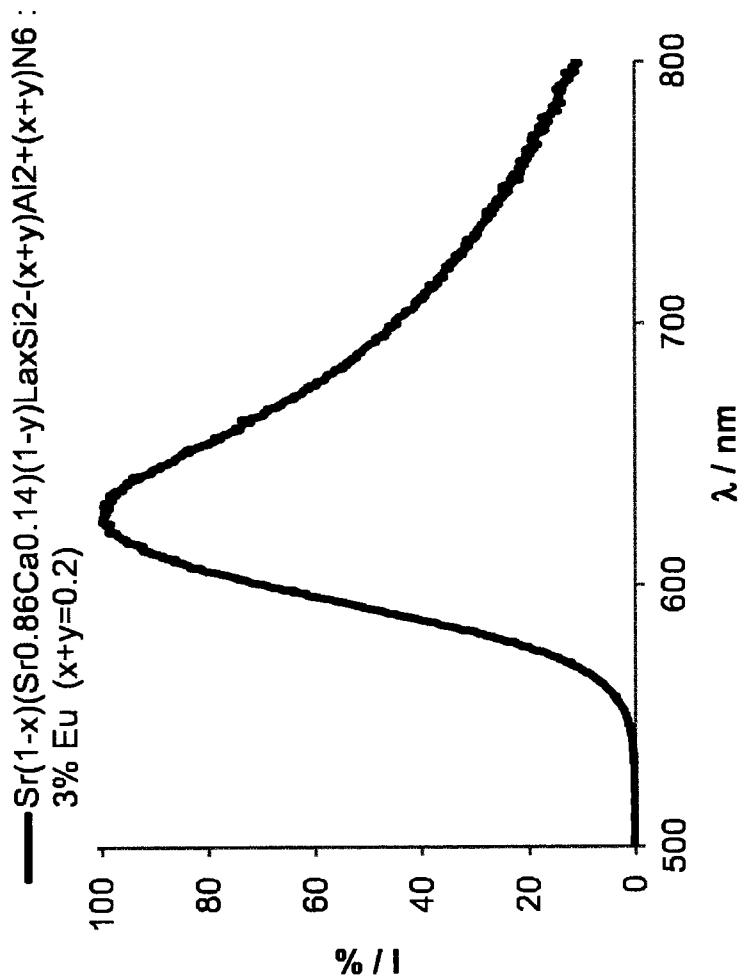


Fig. 59f

Fig. 59g

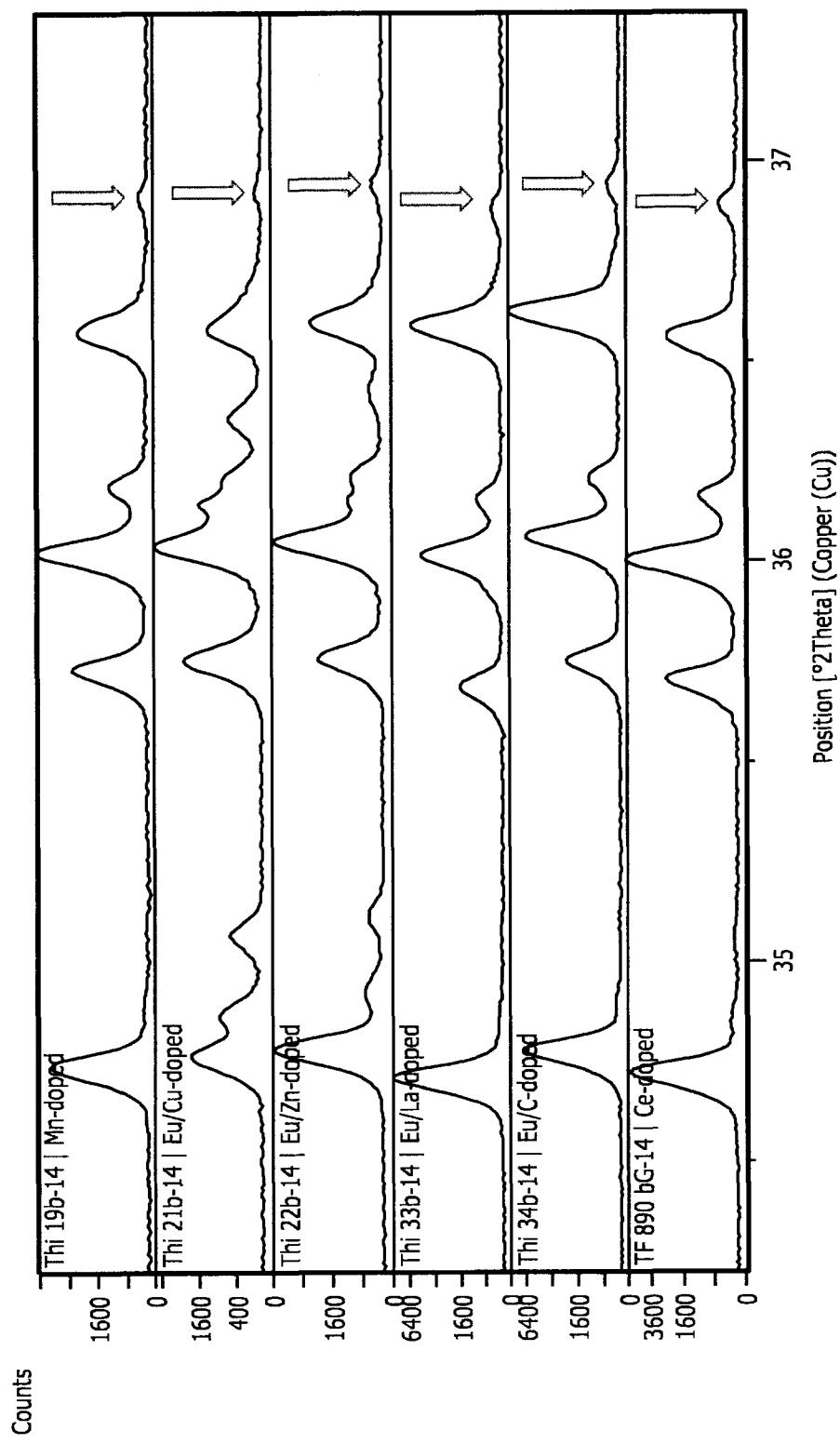


Fig. 60a

Nominal composition	R(450-470)	CIE-x:	CIE-y:	LER	ldom [nm]	Centroid WL [nm]	FWHM [nm]
Sr(Sr _{0,8} Ca _{0,2})Si ₂ Al ₂ N ₆ : 3% Eu	7,2%	0,653	0,346	0,275	606,4	641,1	77,6
Sr(Sr _{0,8} Ca _{0,2})(Si _{1,98} C _{0,02})Al ₂ N ₆ : 3% Eu	7,8%	0,656	0,343	0,258	607,3	643,8	79,8
Sr(Sr _{0,8} Ca _{0,2})(Si _{1,60} C _{0,40})Al ₂ N ₆ : 3% Eu	8,2%	0,656	0,343	0,245	607,3	646,2	83,5

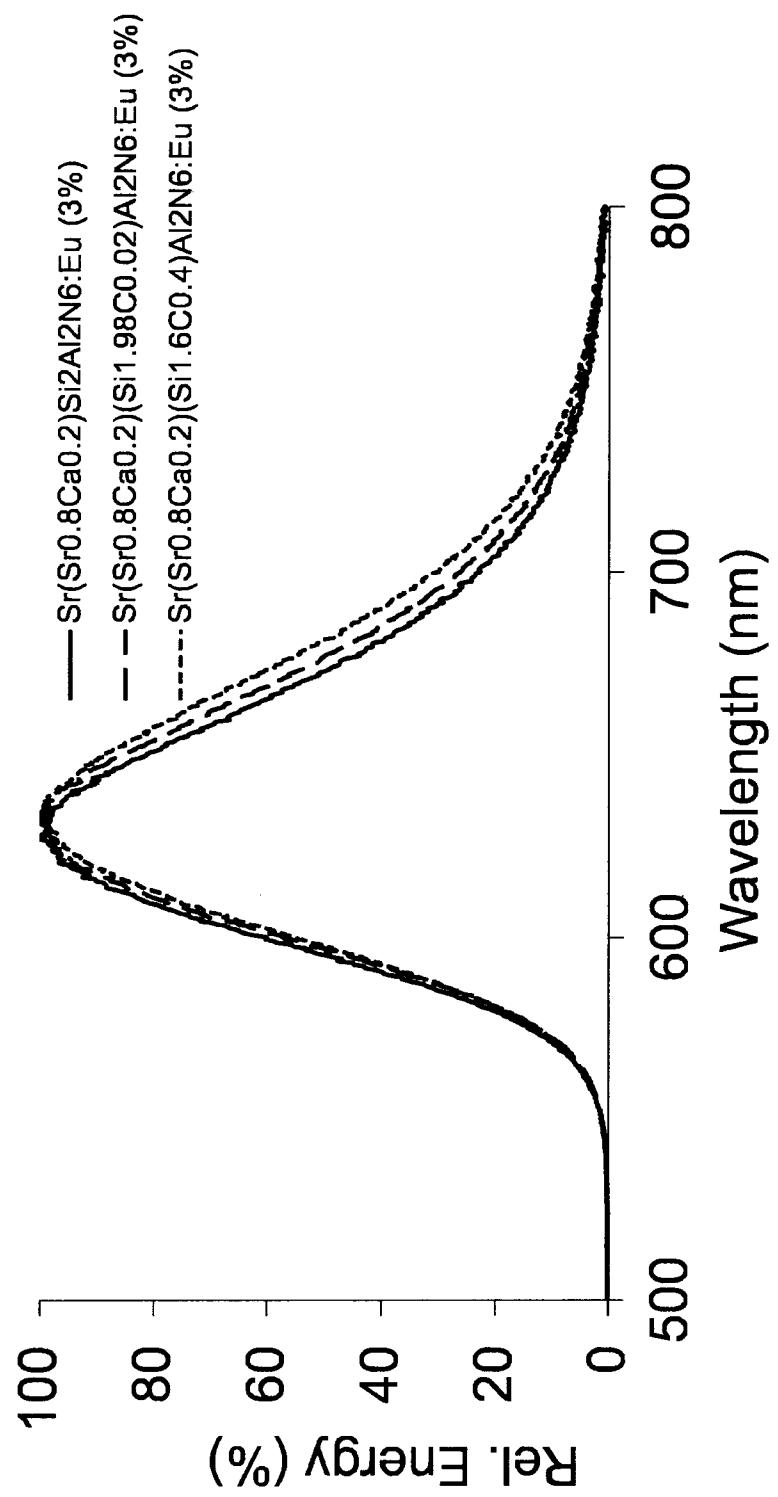


Fig. 60b

Fig. 61a

Nominal composition	R(450-470)	CIE-x:	CIE-y:	LER	λ _{dom} [nm]	FWHM [nm]
Sr(Sr _{0.8} Ca _{0.2})Si ₂ Al ₂ N ₆ :Eu(1,5%),Ce(1,5%),Li(1,5%)	13,3%	0,638	0,36	0,292	602,8	88,0
Sr(Sr _{0.8} Ca _{0.2})Si ₂ Al ₂ N ₆ :Mn(2%)	79,6%	0,573	0,423	0,480	590,1	85,5
Sr(Sr _{0.8} Ca _{0.2})Si ₂ Al ₂ N ₆ :Ce(1%)	45,2%	0,452	0,524	0,611	572,4	126,9

Fig. 61b

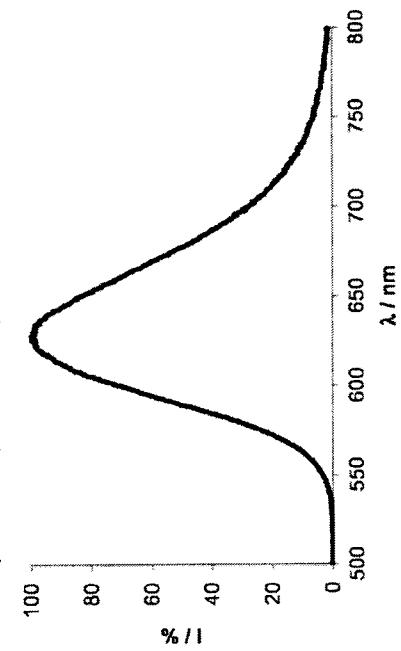


Fig. 61c
Sr(Sr0.8Ca0.2)Si2Al2N6:Mn(2%)

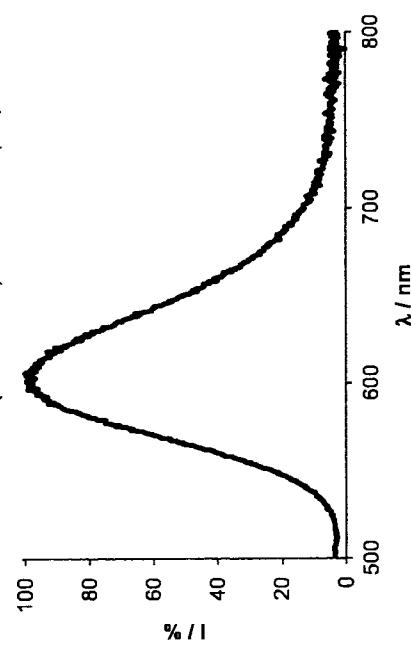
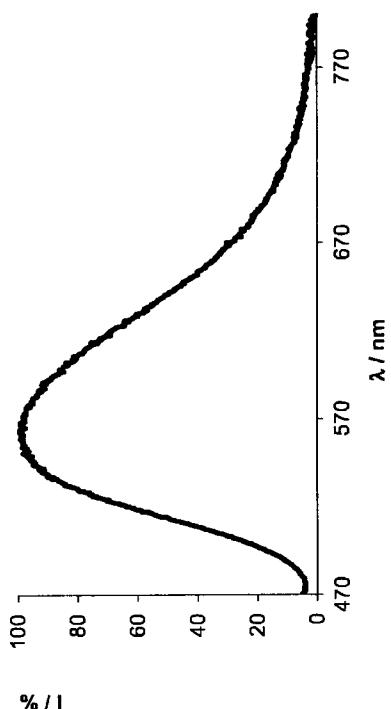


Fig. 61d
Sr(Sr0.8Ca0.2)Si2Al2N6:Ce(1%)



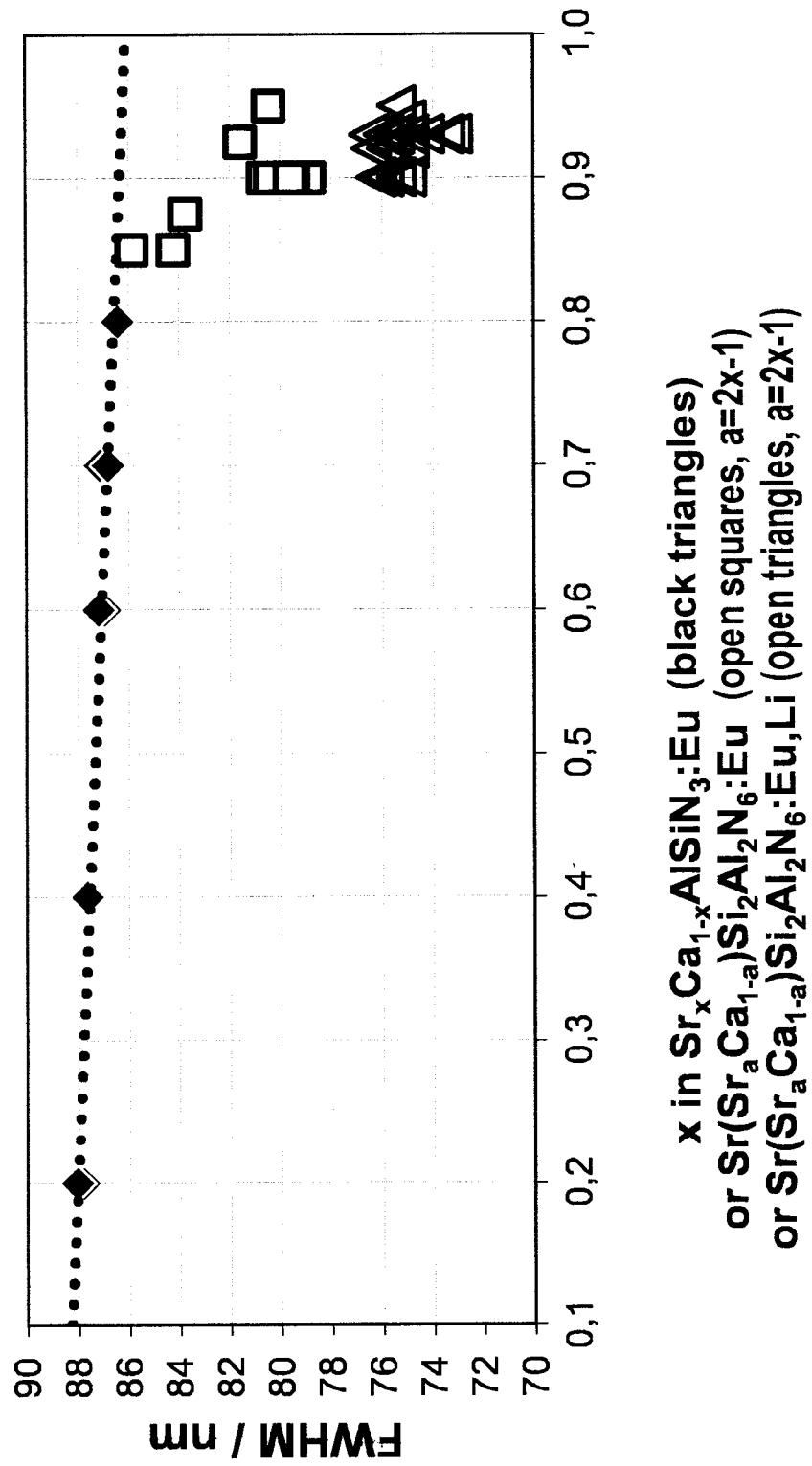
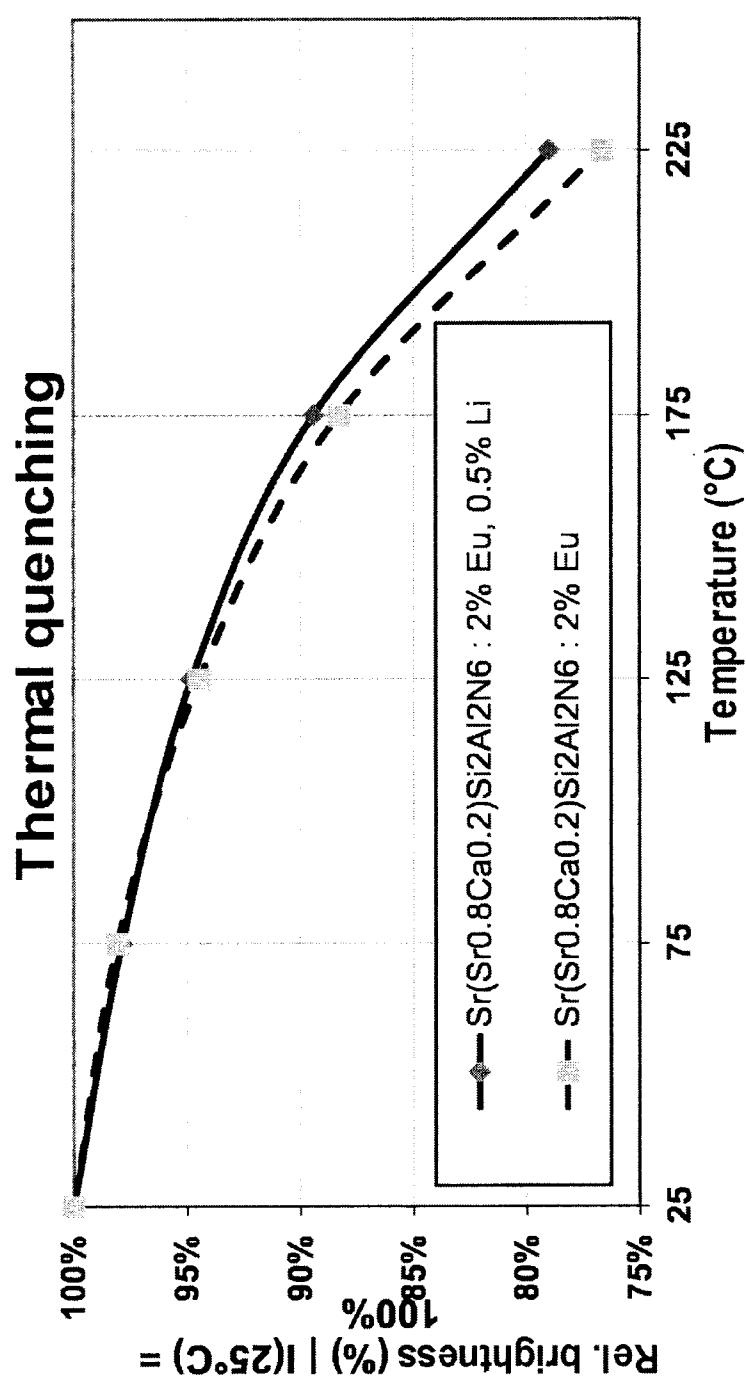


Fig. 62a

Fig. 62b



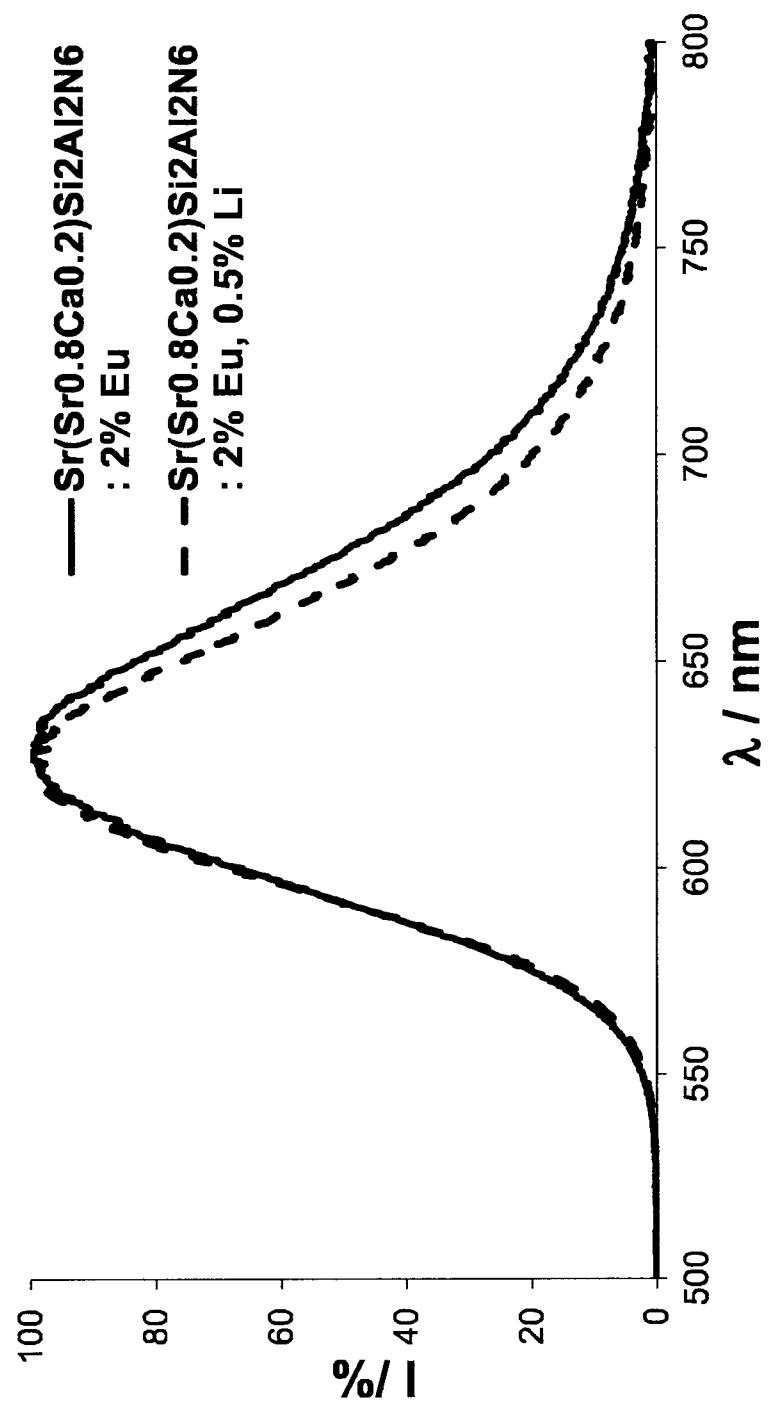


Fig. 62c

Fig. 62d

Composition	rel. Q.E.	CIE-x:	CIE-y:	LER	Centroid WL [nm]	λ_{don} [nm]	FWHM [nm]
$\text{Sr}(\text{Sr}_{0,8}\text{Ca}_{0,2})\text{Si}_2\text{Al}_2\text{N}_6\text{-Eu (2\%)}, \text{Li (0,5\%)}$	107%	0,645	0,353	0,304	637,1	604,6	77,4
$\text{Sr}(\text{Sr}_{0,8}\text{Ca}_{0,2})\text{Si}_2\text{Al}_2\text{N}_6\text{-Eu (2\%)}$	100%	0,644	0,355	0,284	641,0	604,1	84,5

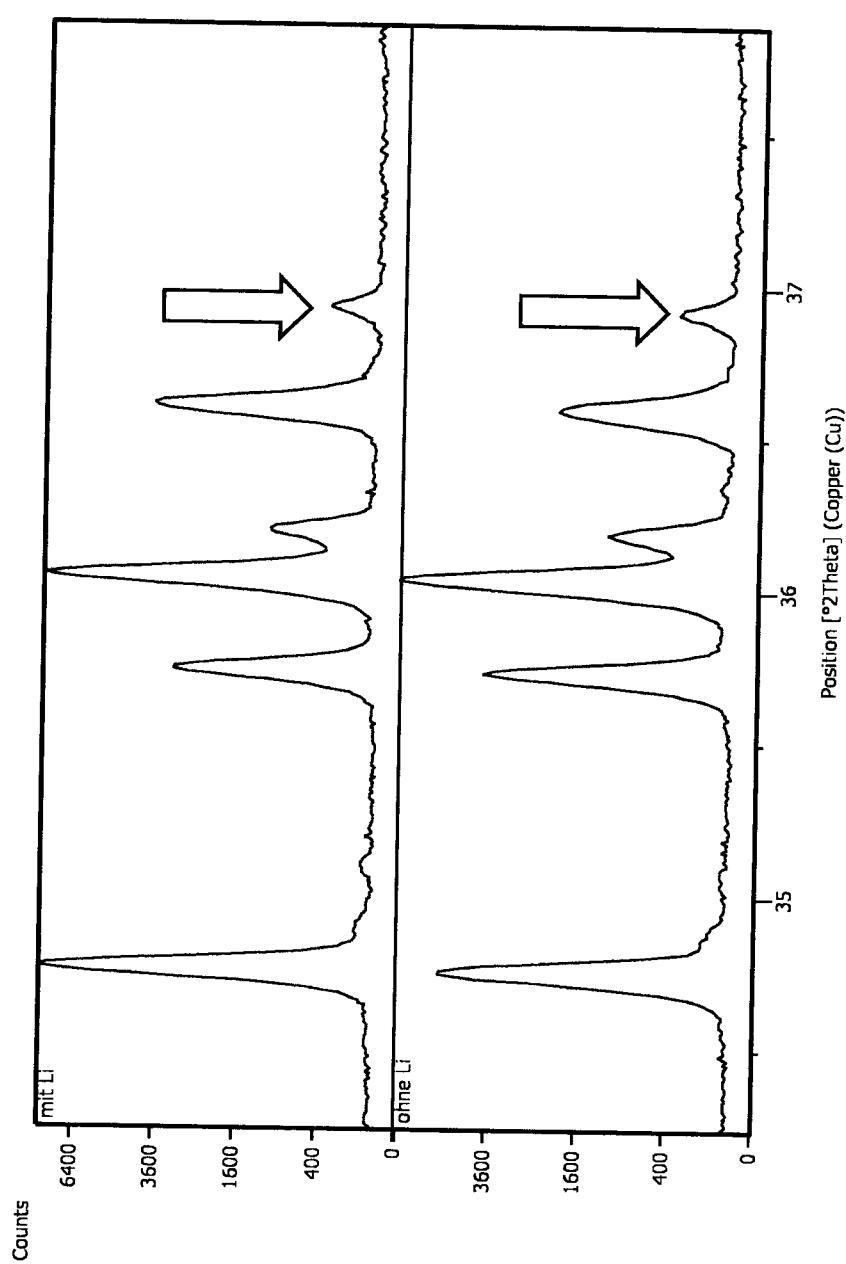


Fig. 62e

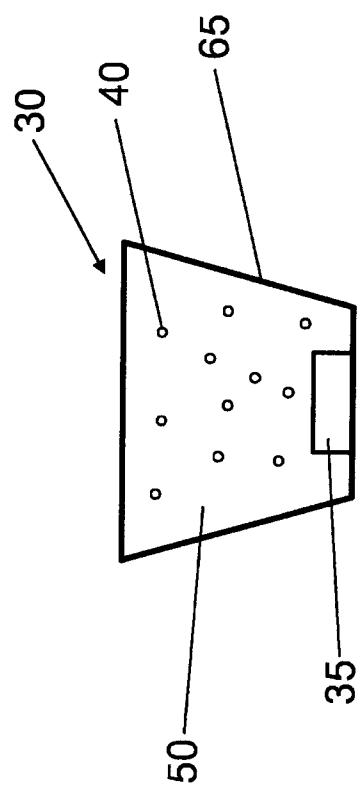


Fig. 63

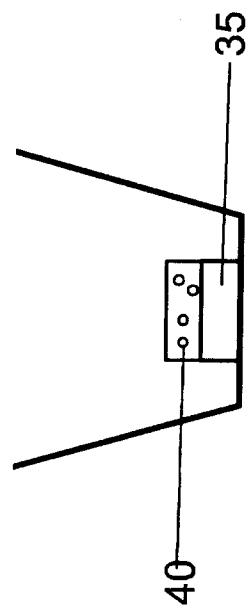


Fig. 64

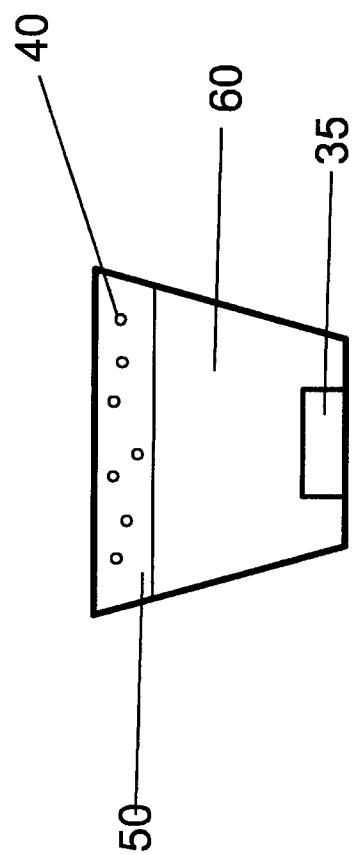


Fig. 65

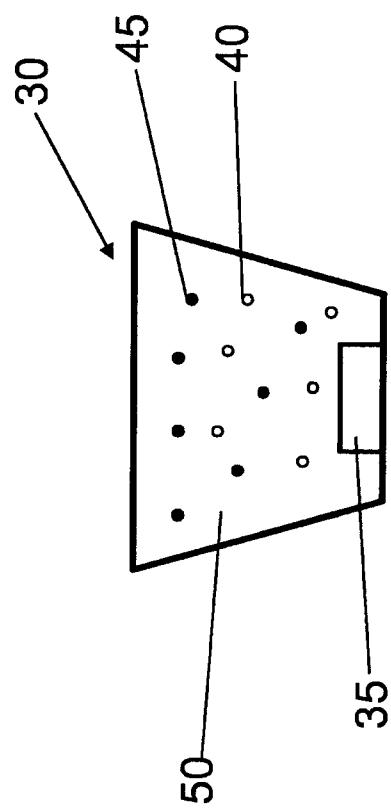


Fig. 66

Fig. 67

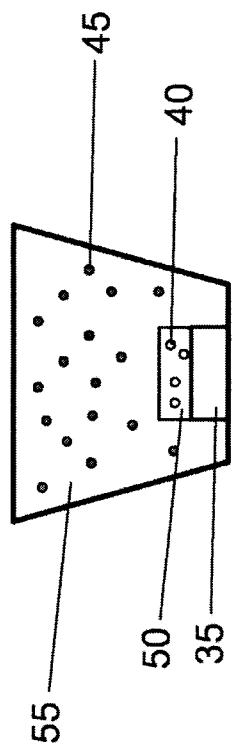
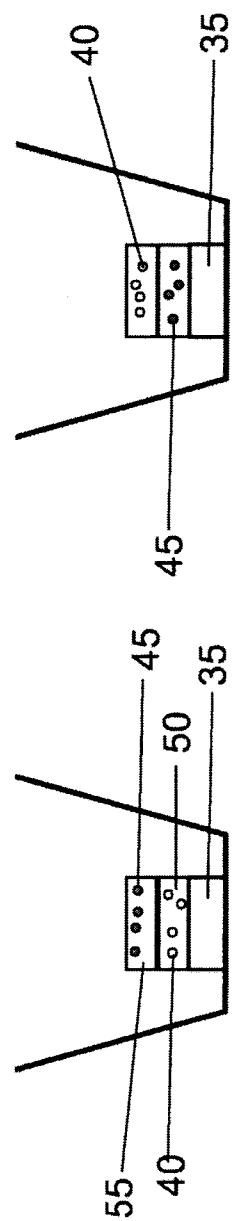


Fig. 68a and 68b



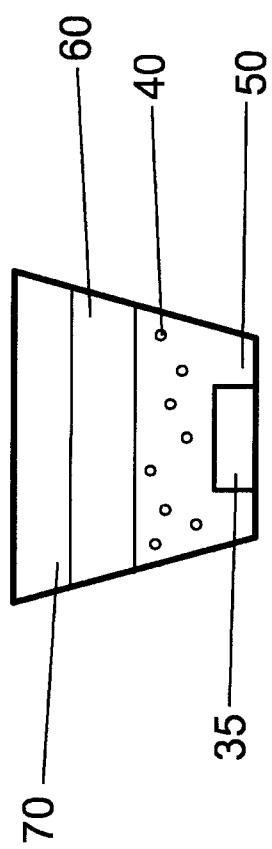


Fig. 69

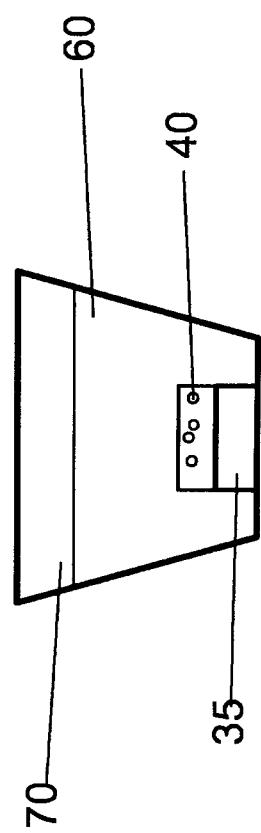


Fig. 70

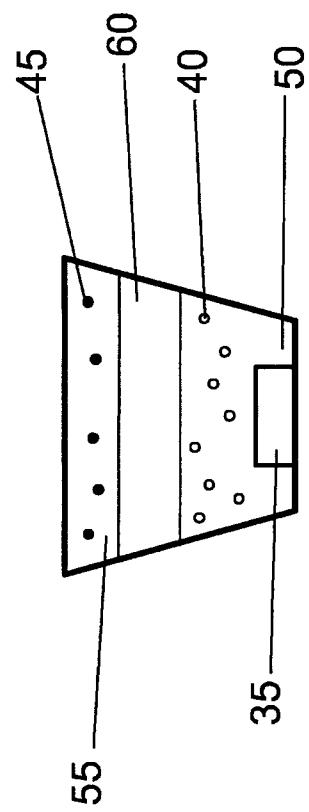


Fig. 71

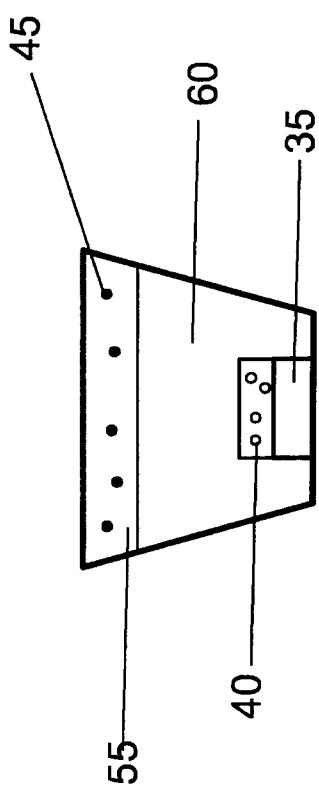


Fig. 72

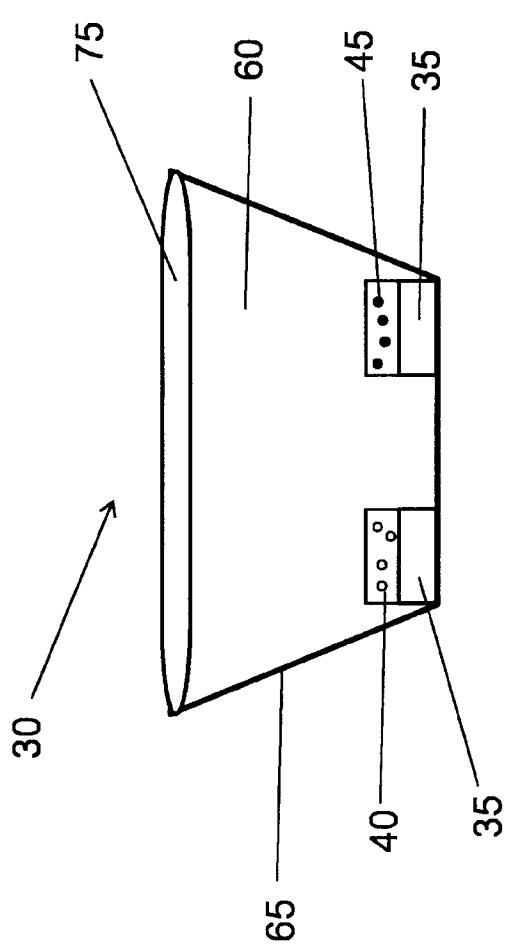


Fig. 73a

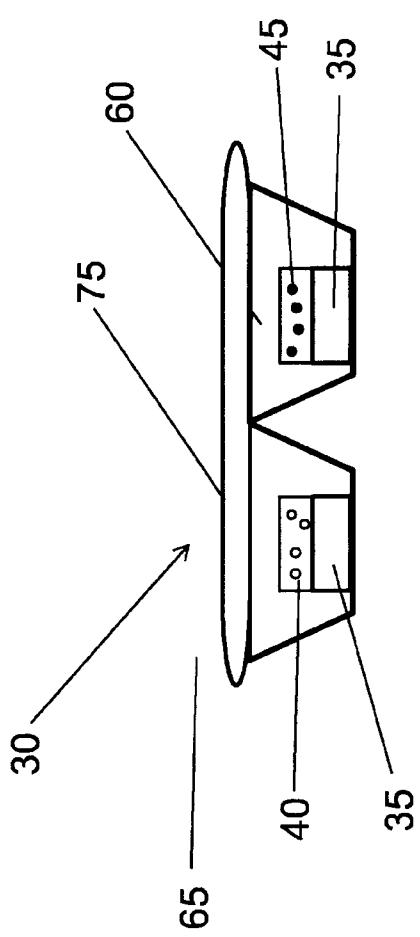


Fig. 73b

Fig. 74

No.	Ca _{1-x} Sr _x AlSiN ₃ -Eu _y										Color locus				
	x	x	Ca ₃ N ₂	Si ₃ N ₂	AlF ₃	SrF ₂	BaF ₂	Li ₂ BaO ₇	LiBF ₄	Si ₃ N ₄	AlN	Eu ₂ O ₃	CIE x	CIE y	
1	0.6%	0.8	2.746	21.552	0.875	0.436	0	0	0	12.993	11.104	0.293	0.632	0.365	
2	0.6%	0.8	2.730	21.428	0.290	1.301	0	0	0	12.918	11.040	0.292	0.632	0.365	
3	0.6%	0.8	2.729	21.416	0.232	1.387	0	0	0	12.911	11.034	0.292	0.634	0.364	
4	0.6%	0.8	2.748	21.564	0.934	0.349	0	0	0	13.000	11.111	0.294	0.632	0.365	
5	0.6%	0.8	2.738	21.492	1.163	0	0	0	0	12.957	11.357	0.293	0.631	0.366	
6	0.6%	0.8	2.646	20.789	1.124	1.682	0	0	0	12.521	10.975	0.283	0.631	0.366	
7	0.6%	0.8	2.574	20.202	1.641	2.454	0	0	0	12.179	10.676	0.275	0.633	0.364	
8	0.6%	0.8	2.506	19.665	2.129	3.185	0	0	0	11.855	10.392	0.268	0.632	0.364	
9	0.6%	0.8	2.731	21.434	0	0	0	0	0	1.295	12.922	11.326	0.292	0.635	
10	0.6%	0.8	2.735	21.463	0.581	0	0	0	0	0.648	12.939	11.342	0.292	0.633	
11	0.6%	0.8	2.719	21.340	0.000	0.864	0	0	0	0.645	12.865	11.277	0.290	0.634	
12	0.6%	0.8	2.725	21.389	0.387	0.574	0	0	0	0.432	12.895	11.303	0.291	0.633	
13	0.6%	0.8	2.671	20.961	0	0	0	0	0	0	12.637	11.077	0.285	0.633	
14	0.6%	0.8	2.704	21.223	0.574	0	1.189	0	0	0	12.795	11.215	0.289	0.631	0.366
15	0.6%	0.8	2.689	21.103	0.000	0.855	1.193	0	0	0	12.722	11.152	0.287	0.630	0.367
16	0.6%	0.8	2.701	21.195	0.000	0	1.198	0	0	0	12.778	11.200	0.289	0.636	0.364
17	0.4%	0.6	5.741	16.896	0.305	1.368	0	0	0	13.581	11.905	0.204	0.638	0.380	
18	0.4%	0.6	5.738	16.886	0.244	1.438	0	0	0	13.573	11.897	0.204	0.635	0.361	
19	0.4%	0.6	5.776	16.999	0.920	0.459	0	0	0	13.664	11.977	0.206	0.637	0.359	
20	0.4%	0.6	5.779	17.009	0.982	0.367	0	0	0	13.672	11.984	0.206	0.638	0.358	
21	0.4%	0.6	5.793	17.048	0.000	0	0	0	0	1.239	13.703	12.011	0.206	0.636	
22	0.4%	0.6	5.793	17.049	0.615	0	0	0	0	0.620	13.704	12.012	0.206	0.640	
23	0.4%	0.6	5.758	16.946	0	0.915	0	0	0	0.616	13.621	11.940	0.205	0.639	
24	0.4%	0.6	5.785	17.025	0	0	0.619	0.686	0	0	13.685	11.995	0.206	0.642	0.354

Fig. 74

25	0.4%	0.6	5,747	16,913	0	1,205	0	0.418	0	13,595	11,917	0.205	0.639	0.358
26	0.4%	0.6	5,737	16,885	0	1,458	0	0.245	0	13,573	11,897	0.204	0.639	0.358
27	0.4%	0.6	5,782	17,017	0	0	0	0.421	0.905	13,679	11,990	0.206	0.640	0.356
28	0.4%	0.6	5,780	17,011	0	0	0	0.247	1,096	13,674	11,986	0.206	0.640	0.356
29	0.4%	0.6	5,793	17,050	0.809	0	0	0.421	0	13,707	12,013	0.206	0.639	0.357
30	0.4%	0.6	5,793	17,049	0.984	0	0	0.248	0	13,707	12,013	0.206	0.639	0.358
31	0.4%	0.6	5,770	16,980	0.408	0,611	0	0.411	0	13,651	11,964	0.205	0.639	0.358
32	0.4%	0.6	5,726	17,196	0.486	0,728	0	0.245	0	13,545	11,871	0.204	0.639	0.358
33	0.4%	0.6	5,704	17,134	0.906	0,606	0	0.122	0	13,496	11,828	0.203	0.636	0.360
34	0.4%	0.6	5,753	17,280	0,000	1,100	0	0.123	0	13,611	11,929	0.205	0.637	0.360
35	0.4%	0.6	5,793	17,049	1,107	0,000	0	0.124	0	13,707	12,013	0.206	0.638	0.358
36	0.4%	0.6	5,722	17,186	0,818	0,548	0	0.122	0	13,537	11,864	0.204	0.638	0.359
37	0.8%	0.8	2,791	21,907	0.119	0	0	0	0	13,209	11,577	0.398	0.634	0.364
38	1.0%	0.8	2,786	21,883	0.118	0	0	0	0	13,183	11,554	0.496	0.638	0.360
39	1.2%	0.8	2,780	21,820	0.118	0	0	0	0	13,157	11,531	0.594	0.641	0.357
40	1.2%	0.9	1,354	23,908	0.115	0	0	0	0	12,814	11,230	0.579	0.633	0.366
41	1.2%	0.875	1,704	23,397	0.116	0	0	0	0	12,898	11,304	0.582	0.636	0.362
42	1.2%	0.85	2,058	22,878	0.117	0	0	0	0	12,983	11,378	0.586	0.640	0.358
43	1.2%	0.825	2,417	22,353	0.117	0	0	0	0	13,069	11,454	0.590	0.642	0.356
44	1.2%	0.925	1,009	24,413	0.114	0	0	0	0	12,731	11,158	0.575	0.628	0.370
45	1.2%	0.95	0,668	24,912	0.114	0	0	0	0	12,649	11,086	0.571	0.626	0.372
46	1.5%	0.8	2,772	21,756	0.118	0	0	0	0	13,118	11,497	0.740	0.645	0.353
47	2.0%	0.8	2,759	21,649	0.117	0	0	0	0	13,053	11,440	0.982	0.654	0.345
48	0.5%	0.7	8,632	39,518	0.244	0	0	0	0	27,227	23,866	0.512	0.635	0.364
49	0.6%	0.8	5,594	43,903	0.238	0	0	0	0	26,468	23,200	0.598	0.638	0.361
50	0.5%	0.7	8,632	39,518	0.244	0	0	0	0	27,227	23,866	0.512	0.632	0.366

Fig. 74

51	0,6%	0,8	5,594	43,903	0,238	0	0	0	26,486	23,200	0,598	0,628	0,371
52	0,5%	0,8	5,600	43,947	0,238	0	0	0	26,494	23,223	0,499	0,624	0,375
53	0,6%	0,9	2,724	48,097	0,231	0	0	0	25,774	22,592	0,582	0,622	0,377
54	0,7%	0,9	2,721	48,050	0,231	0	0	0	25,749	22,570	0,678	0,624	0,375
55	0,8%	0,9	2,719	48,004	0,231	0	0	0	25,724	22,548	0,774	0,626	0,372
56	1,0%	0,9	2,713	47,911	0,231	0	0	0	25,674	22,505	0,966	0,635	0,364
57	1,2%	0,9	2,708	47,818	0,230	0	0	0	25,625	22,461	1,157	0,637	0,362
58	0,6%	0,85	4,140	46,028	0,235	0	0	0	26,116	22,892	0,590	0,624	0,374
59	0,9%	0,9	2,716	47,957	0,231	0	0	0	25,699	22,527	0,870	0,629	0,370
60	1,2%	0,95	1,337	49,826	0,227	0	0	0	25,295	22,173	1,142	0,631	0,368

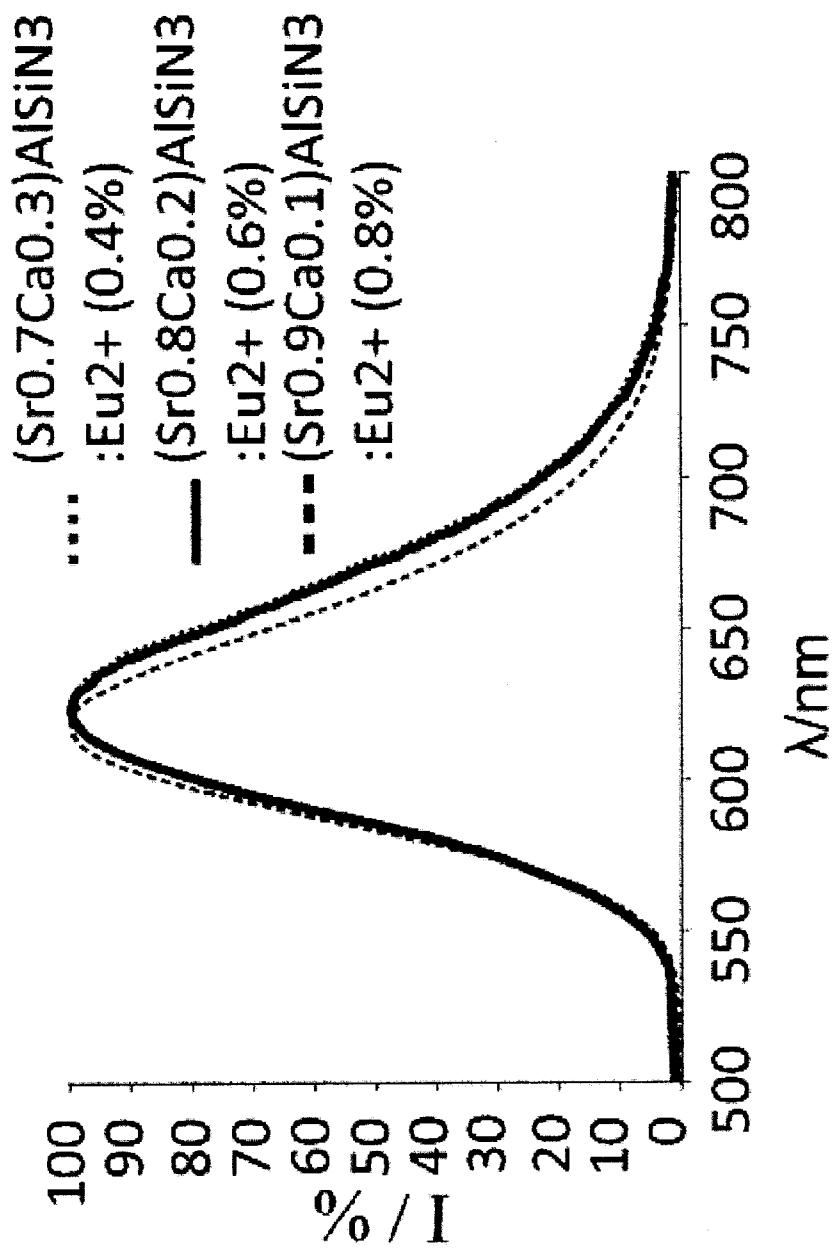


Fig. 75

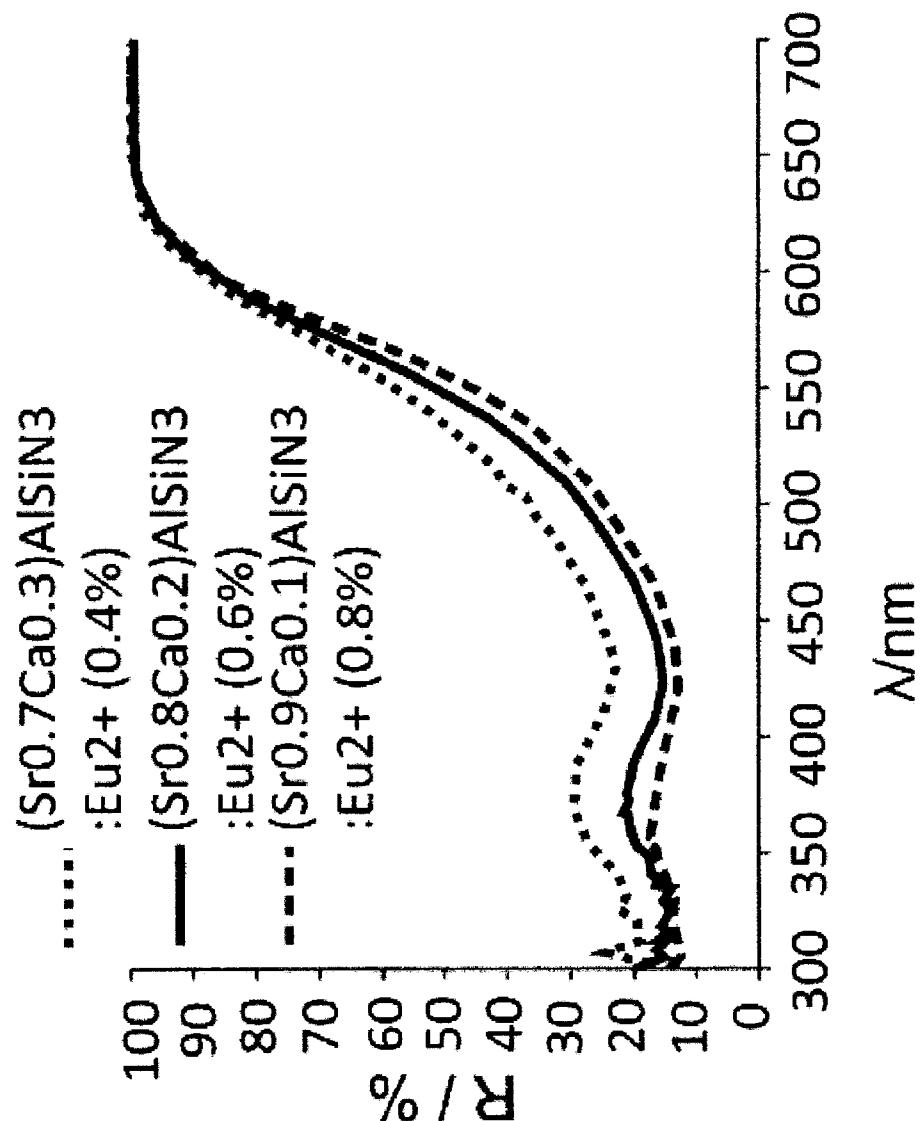


Fig. 76

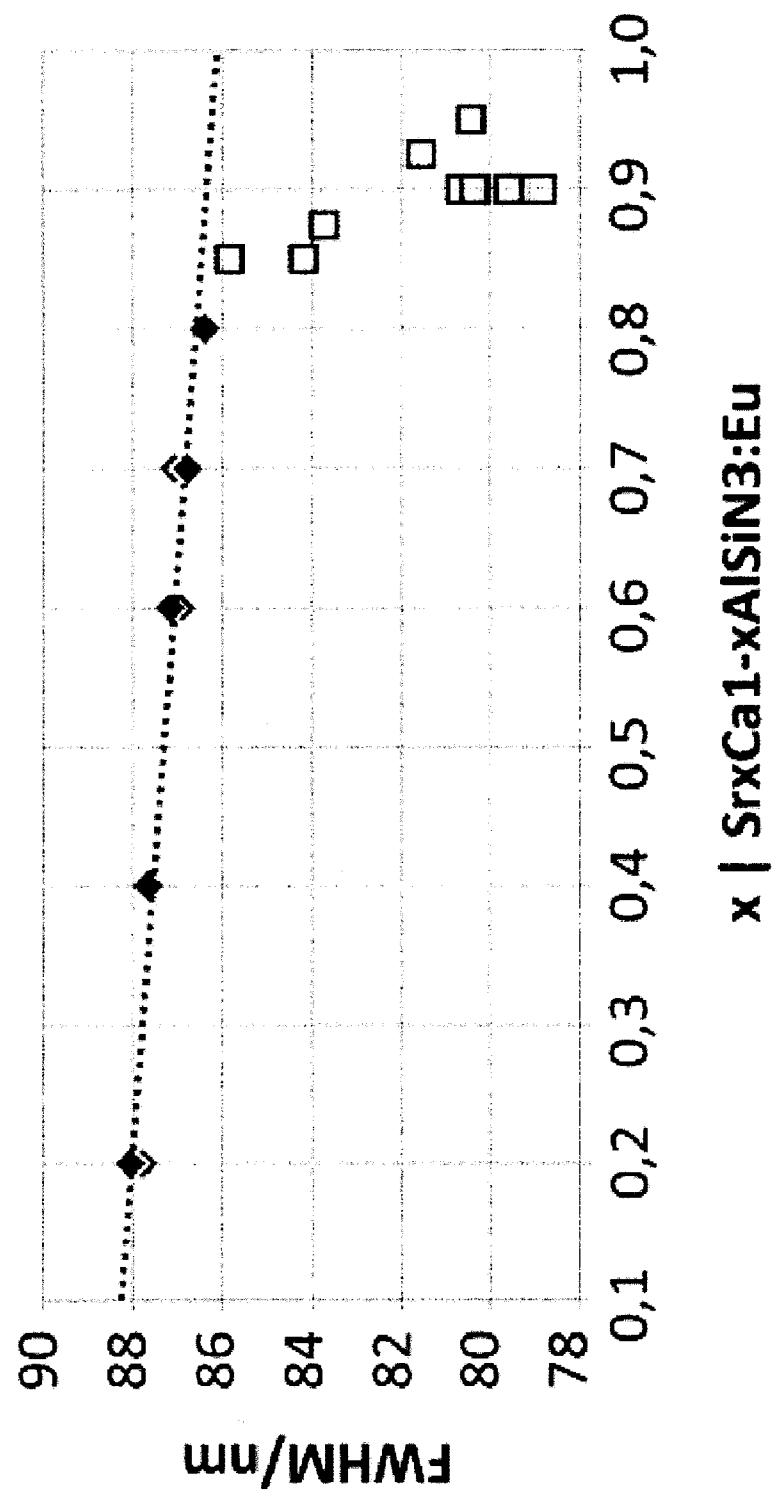


Fig. 77

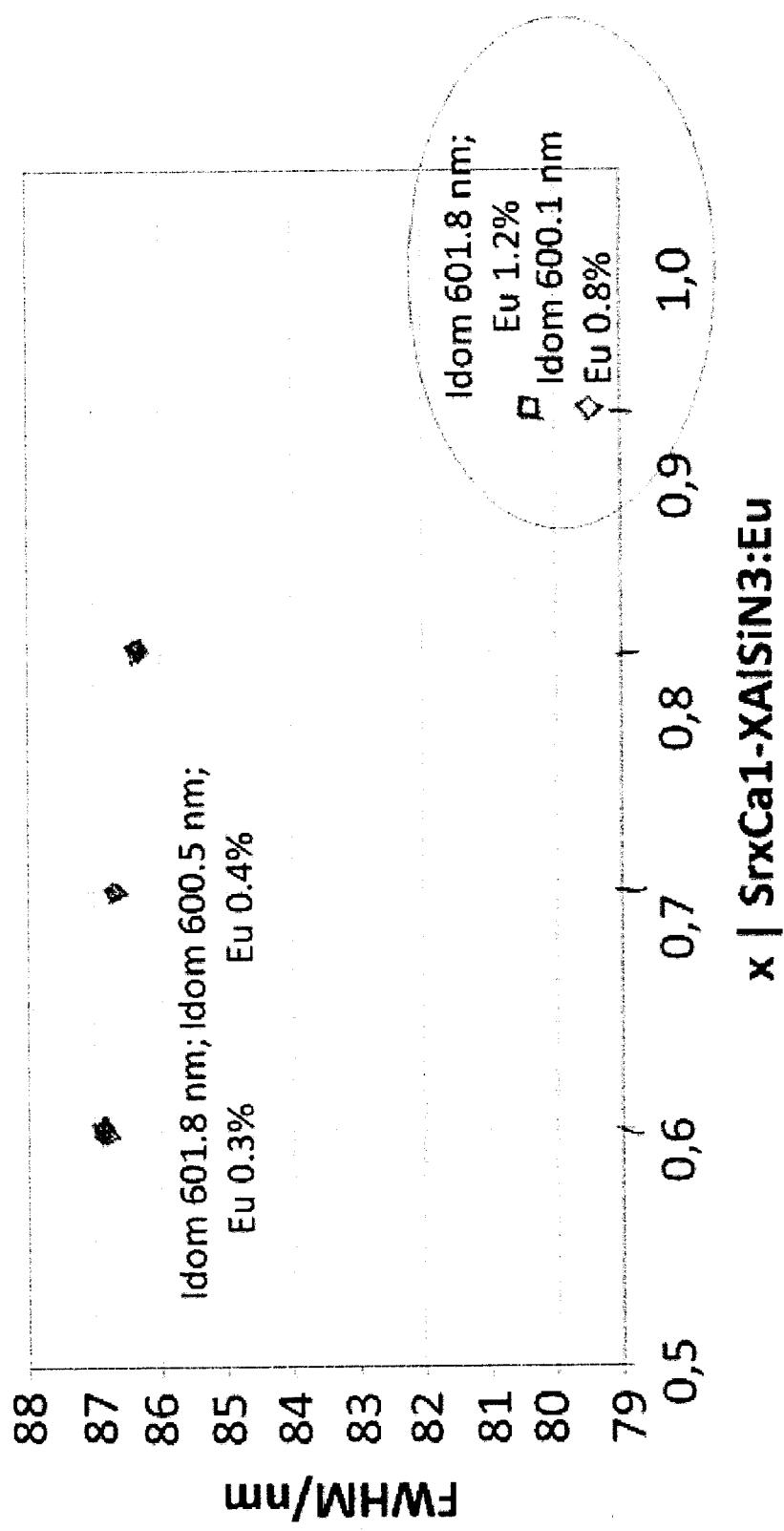


Fig. 78

Fig. 79

	x	y	Idom/nm	FWHM/nm	QI	QE	LER	B
(Sr _{0.7} Ca _{0.3})AlSiN ₃ :Eu (0.4%)	0,628	0,37	600,5	86,7	100%	100%	100%	100%
(Sr _{0.8} Ca _{0.2})AlSiN ₃ :Eu (0.6%)	0,628	0,371	600,3	86,5	99%	110%	101%	111%
(Sr _{0.9} Ca _{0.1})AlSiN ₃ :Eu (0.8%)	0,626	0,372	600,1	79,5	99%	113%	110%	125%

Fig. 80

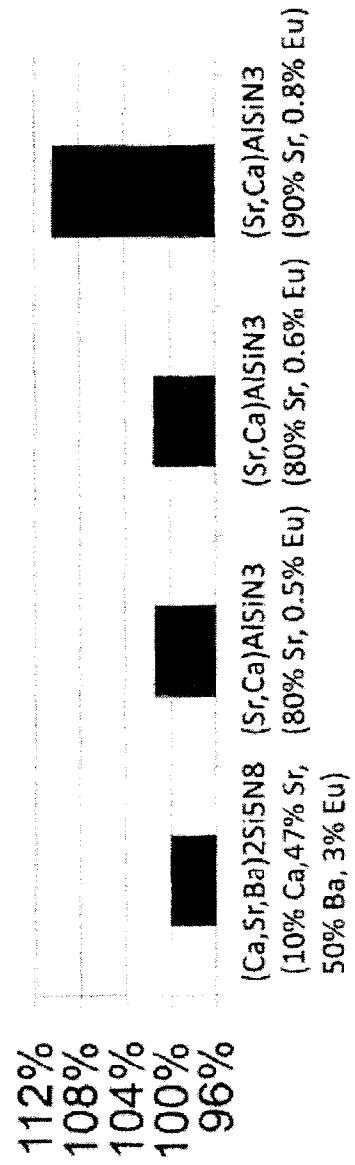


Fig. 81

G	R	(Ca,Sr,Ba)2Si5N8 (10% Ca, 47% Sr, 50% Ba, 3% Eu)	Idom/nm	c	v	CIE x	CIE y	CCT/K	CRI	E
(Lu,Ce)3(Al,Ga)5012 (25% Ga, 2.5% Ce)	(Sr,Ca)AlSiN3 (80% Sr, 0.5% Eu)	601,5	15,50%	4,45:1	0,461	0,408	2669	79	100,00%	
(Lu,Ce)3(Al,Ga)5012 (25% Ga, 2.5% Ce)	(Sr,Ca)AlSiN3 (80% Sr, 0.6% Eu)	599,4	18,00%	2,5:1	0,459	0,416	2754	80	101,40%	
(Lu,Ce)3(Al,Ga)5012 (25% Ga, 2.5% Ce)	(Sr,Ca)AlSiN3 (90% Sr, 0.8% Eu)	600,1	16,70%	3,7:1	0,462	0,411	2681	79	110,40%	

Fig. 82

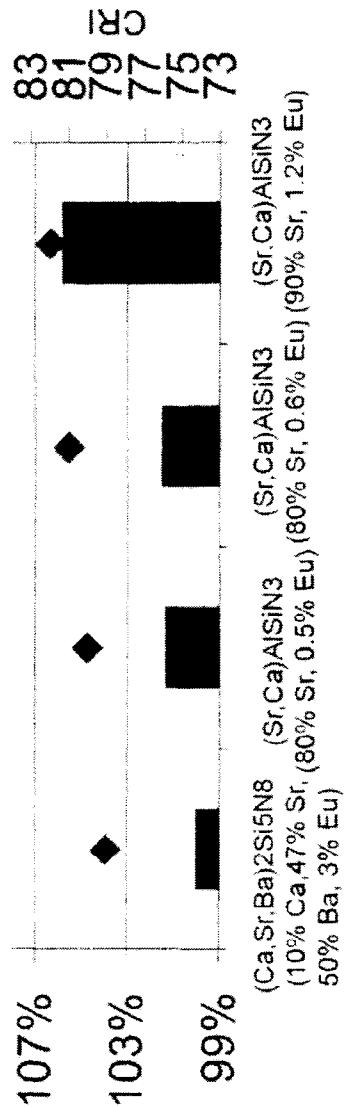
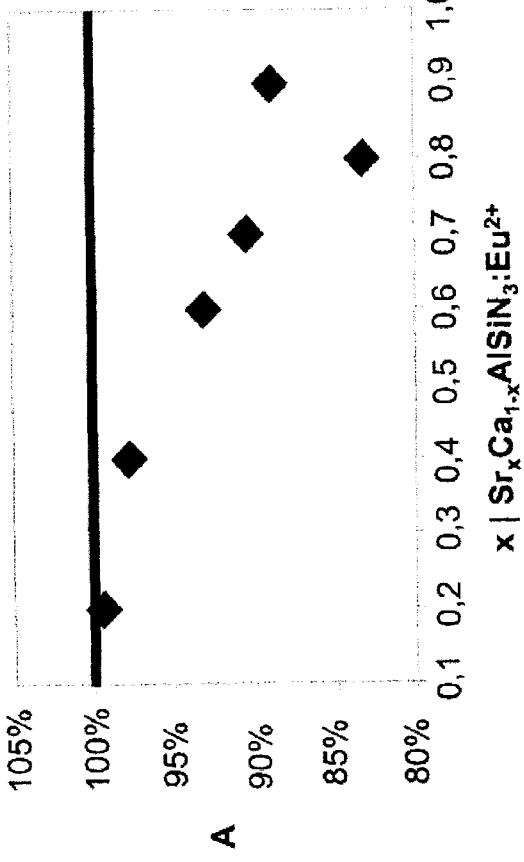


Fig. 83

G	R	Idom/nm	C	V	CIE x	CIE y	CCT/K	CRI	E
(Lu,Ce)3(Al,Ga)5O12	(Ca,Sr,Ba)2Si5N8								
(25% Ga, 2.5% Ce)	(10% Ca, 47% Sr, 50% Ba, 3% Eu)	601,5	15,50%	4,45:1	0,461	0,408	2669	79	100,00%
(Lu,Ce)3(Al,Ga)5O12	(Sr,Ca)AlSiN ₃ (80% Sr, 0.5% Eu)	599,4	18,00%	2,5:1	0,459	0,416	2754	80	101,40%
(25% Ga, 2.5% Ce)	(Sr,Ca)AlSiN ₃ (80% Sr, 0.6% Eu)	600,3	16,80%	2,6:1	0,461	0,408	2655	81	101,50%
(Lu,Ce)3(Al,Ga)5O12	(Sr,Ca)AlSiN ₃ (90% Sr, 1.2% Eu)	602,4	16,50%	5,5:1	0,462	0,408	2649	82	105,80%

Fig. 84



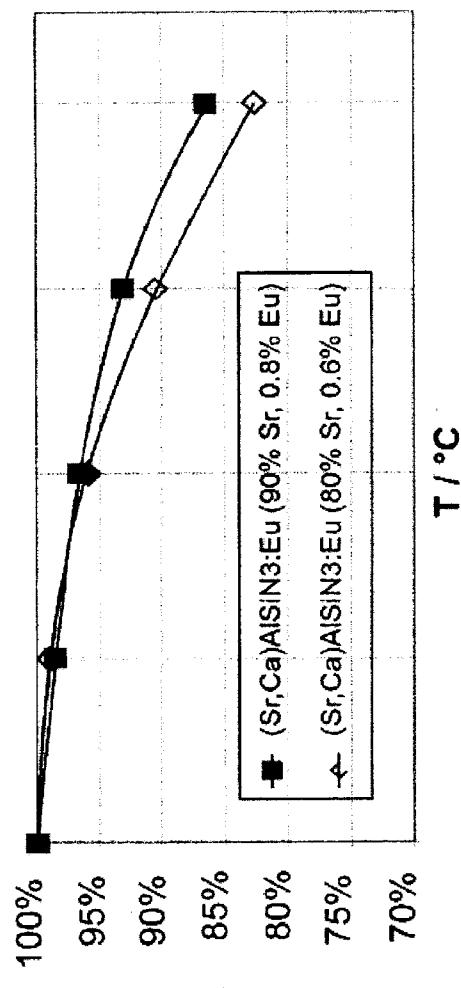


Fig. 85

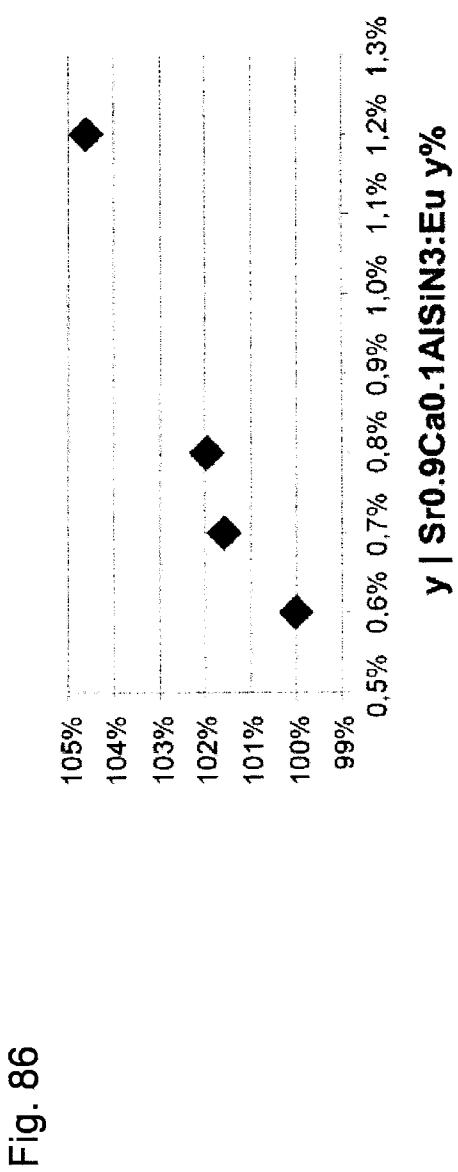
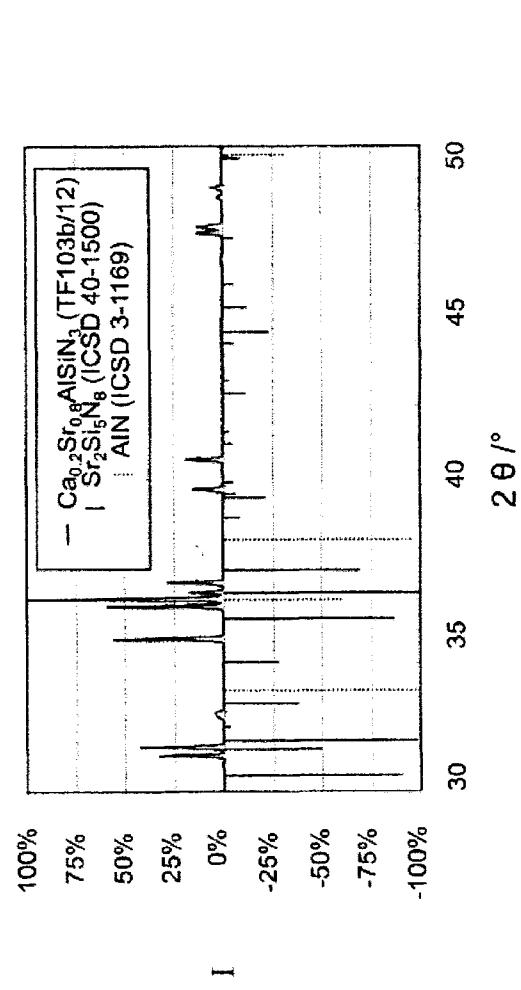
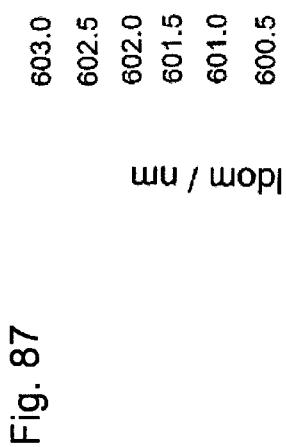


Fig. 86



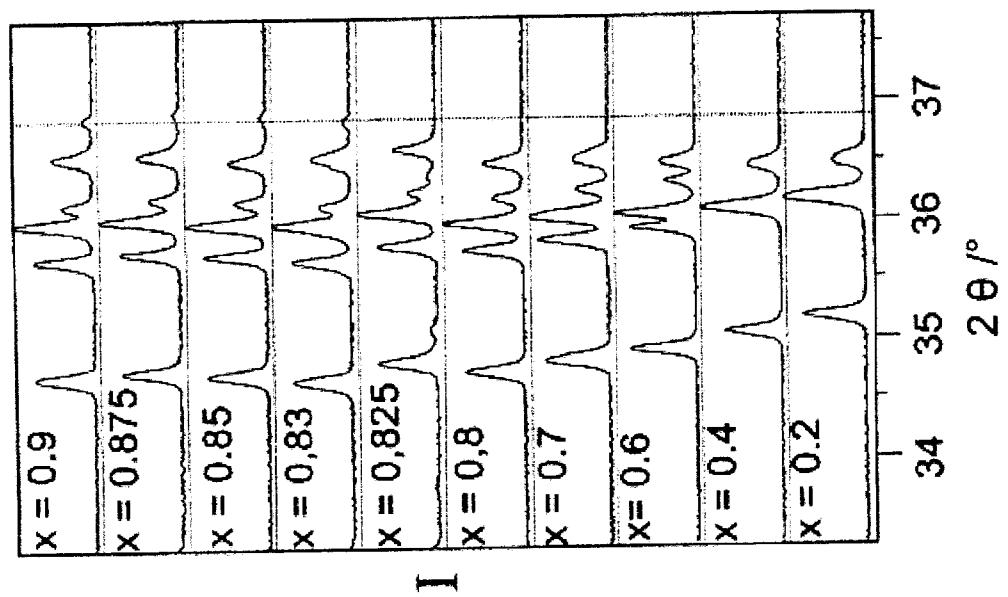


Fig. 89

Space group	Cmc2 ₁	Cell search	P1 (CA only in one layer, unrefined)	P1 final refining
Lattice parameters				
a	9.821 Å	9.825 Å	9.821 Å	9.821 Å
b	5.749 Å	5.752 Å	5.749 Å	5.749 Å
c	5.163 Å	5.165 Å	5.163 Å	5.163 Å
V			291.5 Å ³	291.5 Å ³
Goodness parameters				
R_{wp}	7.2%		7.1 %	6.2%
R_{bragg}	11.1%		11.0 %	9.6%
	5.6%		5.6 %	3.7%

Fig. 90

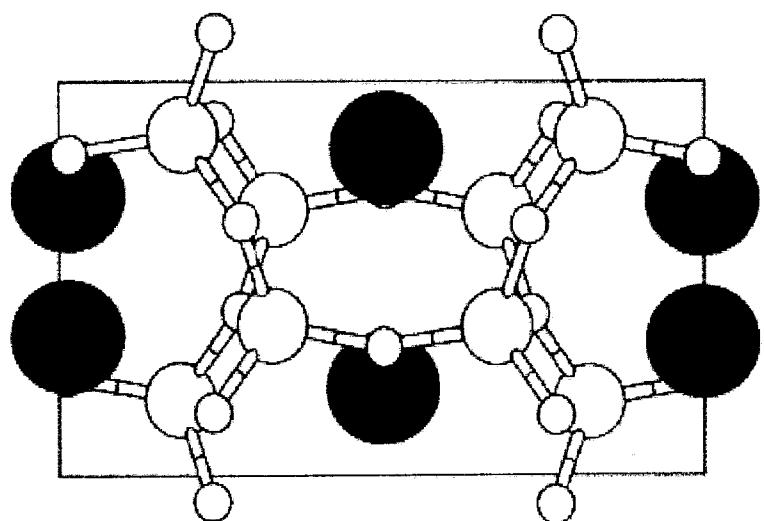


Fig. 92

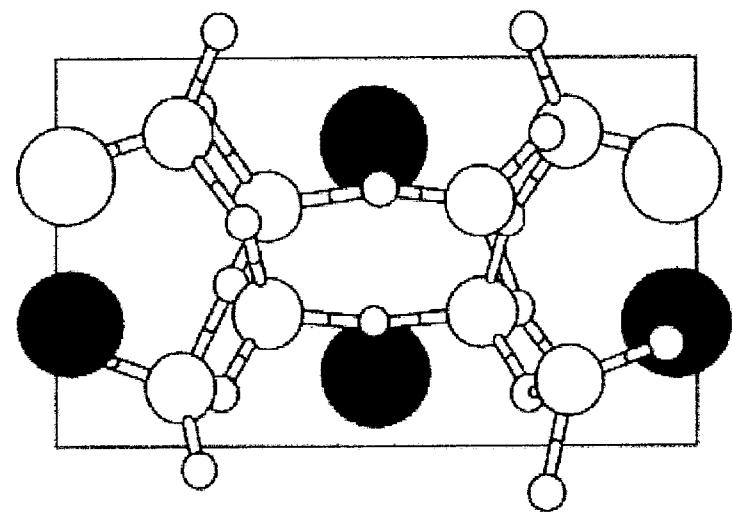


Fig. 91

