



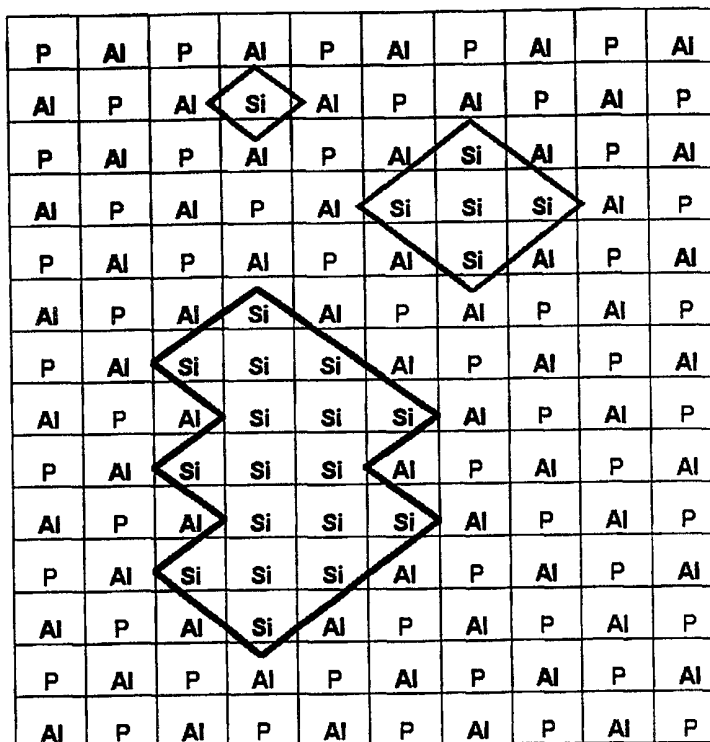
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(54) Title: NEW SILICOALUMINOPHOSPHATES HAVING AN AEL STRUCTURE, AND THEIR PREPARATION

(57) Abstract

Disclosed are silicoaluminophosphates (SAPOs) having unique silicon distributions and their preparation. More particularly, the new SAPOs have a high silica:alumina ratio, and are prepared from single phase synthesis solutions, or from microemulsions containing surfactants.



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NEW SILICOALUMINOPHOSPHATES HAVING AN AEL STRUCTURE, AND THEIR PREPARATION

FIELD OF THE INVENTION

This invention relates to silicoaluminates (SAPOs) having unique silicon distributions, a method for their preparation and their use as catalysts. More particularly, the new SAPOs, designated ECR-42 herein, have a high silica:alumina ratio and a beneficial silicon atom distribution.

BACKGROUND OF THE INVENTION

The preparation of crystalline silicoaluminophosphates is well known. U.S. Patent No. 4,480,871 describes the preparation of crystalline, microporous silicoaluminophosphates by hydrothermal crystallization of silicoaluminophosphate gels containing a molecular structure forming template. SAPOs are members of a class known as non-zeolitic molecular sieves. The SAPO molecular sieve has a framework of AlO_4 , SiO_4 and PO_4 tetrahedra linked by oxygen atoms. The negative charge in the network is balanced by the inclusion of exchangeable protons or cations such as alkali or alkaline earth metal ions. The interstitial spaces or channels formed by the crystalline network enables SAPOs to be used as molecular sieves in a manner similar to crystalline aluminosilicates such as zeolites.

Accordingly, numerous microporous framework structures analogous to the aluminosilicate zeolites can be synthesized having an AlPO_4 composition and have been called ALPOs. A modified family of materials has been made by the substitution of Si^{4+} for Al^{3+} and P^{5+} (SAPOs). Although the ALPO structures are neutral frameworks, the substitution of Si^{4+} for P^{5+} imparts a negative charge on the framework. By suitable choice of a cation, this can be translated into catalytic activity. However, alternate substitutions may be

possible that may result in a disproportionately low exchange capacity. The exact nature of Si substitution into ALPO structures is complex and highly variable and may depend on both the topology of the ALPO/SAPO and the method of preparation. The result is that preferred catalysts may be made by a suitable choice of synthesis method. For example, SAPO-5 and SAPO-11 may be conventionally prepared in an aqueous solution or from microemulsions. The latter processes use hexanol and a cationic or neutral surfactant to a two-phase gel leading to the formation of a microemulsion.

The microemulsion process is a two-phase approach to preparing SAPOs attempts to reduce the amount of undesirable silica island formation by supplying the silicon from an organic phase to the aqueous phase at a low concentration during crystallization. The organic phase contains the organic solvent and organic silicon source, tetraethylorthosilicate, which is only slightly soluble in the aqueous phase. The aqueous phase is where crystallization occurs and contains the phosphorous and aluminum. It has been theorized that as the silicon is depleted from the aqueous phases by the growing SAPO crystals, it will be replenished from the organic phase, thereby forming a silicoaluminophosphate product having a more uniform distribution of silicon in the framework.

SAPOs have application for a wide variety of uses, for example as catalysts. In this regard, it is known that increasing Si concentration at first results in an increase in catalytic activity. However, increasing Si content beyond about 0.04 mole fraction in the framework, based on the total amount of silicon, aluminum, and phosphorous in the framework, provides no increase in activity, and may even lead to a decrease, depending on the specific distribution and clustering of the Si^{4+} substituent.

In that the distribution of Si in the SAPO framework affects catalytic activity, the catalytic activity of SAPOs therefore depends on both the

global composition and the Si distribution. Accordingly, SAPOs are defined not only by chemical composition and X-Ray Diffraction, but also by ^{29}Si MAS NMR spectra which define the Si distributions. On the basis of this last technique, it has been shown that when the SAPOs contain low amounts of Si, the silicon atoms are mostly isolated. However, when the Si content increases, Si islands start to appear, i.e., Si sites having silicon atoms and no aluminum atoms in neighboring lattice positions.

It would, therefore, be desirable to further increase silicon content in SAPO molecular sieves without forming undesirable silica islands in order to increase catalytic activity and selectivity.

SUMMARY OF THE INVENTION

In one embodiment, the invention is a composition comprising a molecular sieve with a framework tetrahedra of silicon, aluminum, and phosphorus, and designated ECR-42, the composition having a total silicon amount ranging from above about 4 molar percent to about 20 molar percent, a total aluminum amount ranging from about 40 molar percent to about 55 molar percent, and a total phosphorus amount ranging from about 30 molar percent to about 50 molar percent, the molar percents being based on the total amount of aluminum, phosphorus, and silicon present in the composition, and the molecular sieve having the topology AEL and being isostructural with conventional SAPO-11, wherein

(a) the silicon present in the molecular sieve and the conventional SAPO-11 is distributed among silicon sites in the framework tetrahedra, each site having a first, a second, a third, and a fourth nearest neighbor position, and each next nearest neighbor position being independently occupied by one atom selected from silicon, and aluminum, and

(b) the molecular sieve has a substantially smaller number of silicon sites having silicon atoms among all four nearest neighbor positions than the conventional SAPO-11 with the total silicon amount.

In another embodiment, the invention is a composition comprising a molecular sieve with a framework tetrahedra of silicon, aluminum, and phosphorus, the composition having a total silicon amount ranging from above about 4 molar percent to about 20 molar percent, a total aluminum amount ranging from about 40 molar percent to about 55 molar percent, and a total phosphorus amount ranging from about 30 molar percent to about 50 molar percent, the molar percents being based on the total amount of aluminum, phosphorus, and silicon present in the composition, and the molecular sieve having the topology AEL and being isostructural with conventional SAPO-11, wherein

- (a) the molecular sieve has a first number of Si atoms coordinated as Si (4Si),
- (b) the conventional SAPO-11 has a second number of Si atoms coordinated as Si(4Si), and
- (c) the first number of Si atoms is substantially less than the second number of Si atoms.

In another embodiment, the invention is a method for forming a molecular sieve composition, the method comprising:

- (a) combining at least one water-dispersible phosphorus source material, at least one template capable of structure directing to AEL, water, and a water-soluble or dispersible organic silicon source material under gellation conditions in order to form a gel, and then
- (b) reacting the gel under molecular sieve synthesis conditions.

In another embodiment, the invention is a product formed in accordance with such a process

In another embodiment, the invention is a silicoaluminophosphate gel formed by combining under gellation conditions at least one water-dispersible silicon source material, at least one template capable of structure-directing to AEL, at least one water-dispersible phosphorus source material, and water.

BRIEF DESCRIPTION OF THE DRAWINGS

Figure 1 schematically shows the type of silicon environments in SAPO-type molecular sieves.

Figure 2 shows a representative deconvolution of an ^{29}Si NMR spectrum.

Figure 3 shows the ^{29}Si NMR spectra of conventionally prepared SAPO-11. Spectrum 3(a) shows a SAPO-11 of 5 molar percent Si having about 1/3 of the Si atoms located in silicon islands. Spectrum 3(b) is a SAPO-11 of 4 molar % Si having at most a small amount of Si located in Si islands, and spectrum 3(c) shows a SAPO-11 of 14 molar % Si having extensive Si island formation.

Figure 4 shows powder x-ray diffraction data for the AEL-type SAPOs prepared in examples 1 through 4.

Figure 5-shows powder x-ray diffraction data from AEL-type SAPOs of examples 5 through 8.

Figure 6(a) shows scanning electron micrographs of an AEL-type SAPO prepared in accordance with the method of Example (4), and figure 6(b) shows an electron micrograph of conventionally-prepared SAPO-11.

Figure 7 shows the ^{29}Si NMR spectra of ECR-42 having framework silicon in amounts ranging from about 5 molar % to about 13 molar %. Spectrum 7(a) shows ECR-42 with about 13 molar percent Si and having about 9% of the Si atoms located in silicon islands, representing a substantially reduced amount of framework Si in islands compared to conventional SAPO-11 of about the same Si content (Figure 3(c)). Spectrum 7(b) is of ECR-42 of example 6 containing about 7 molar % Si and having about 10% of the Si located in Si islands, and spectrum 7(c) shows an ECR-42 of example 1 containing about 4 molar % Si having about 4% of the Si located in Si islands.

DETAILED DESCRIPTION OF THE INVENTION

The invention is based in part on the discovery that SAPO materials having the AEL topology and designated ECR-42 herein may be prepared with high silicon concentration and without undesirable silicon island formation. The invention is also based on the discovery that a SAPO precursor gel such as a silicoaluminophosphate gel may be formed from an aqueous synthesis solution containing no added alcohol or surfactant provided the co-solvent is a soluble species capable of maintaining a high dispersion of silica in a synthesis solution. While not wishing to be bound by any theory or model, the co-solvent is believed to inhibit polymerization of the highly dispersed silica species. The preferred methods therefore provide a SAPO molecular sieve containing Si in the tetrahedral framework at a concentration above about 0.04 molar fraction (based on the total amount of aluminum, silicon, and phosphorous in the framework) and having a desirable Si distribution in the framework, resulting in high catalytic activity and selectivity.

In other words, it has been discovered that modifying the synthesis of a silicon-substituted aluminophosphate by changing the composition of the synthesis mixture, the synthesis conditions, or both, modifies the silicon distribution in the silicoaluminophosphate thus formed. This changed

distribution of silicon may have a major beneficial influence on the catalytic activity of the silicoaluminophosphate.

Accordingly, when AEL-type molecular sieve materials are synthesized following the procedure described herein, the distribution of Si and therefore the total number and strength of acid sites in the molecular sieve framework is quite different, and much higher than those of previously reported forms of SAPO molecular sieves.

The preferred silicoaluminophosphate composition of this invention, ECR-42, has the topology AEL which corresponds to SAPO-11. The AEL topology is defined in the "*Atlas of Zeolite Structure Types*," 4th Ed, by W.M. Meier, D.H. Olson and Ch. Baerlocher, Elsevier, 1996. Although the composition is isostructural with other AEL molecular sieve materials, it is a distinct molecular sieve composition because the silicon, aluminum, and phosphorus atoms are not arranged the same way as in conventional SAPO-11 molecular sieve. Those skilled in the art will recognize that two isostructural molecular sieves may be entirely different compositions having entirely different properties, depending on, for example, the nature and distribution of the constituent atoms.

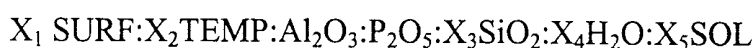
Preferred molecular sieve compositions are physically different from other SAPO's having the AEL structure because the silicon atoms are distributed differently in the molecular sieve framework. The physical structure of the preferred composition (and its silicon distribution) is illustrated schematically in figure 1. While the actual structure is three dimensional and contains oxygen in addition to silicon, aluminum and phosphorus, the figure's atomic positions are represented on a two dimensional array and oxygen atoms are omitted for clarity. As is shown in the figure, each lattice site in the framework has four nearest neighbor lattice sites. In the preferred compositions, as with all AEL-type SAPOs, a lattice site occupied by a silicon 4+ cation, i.e., a

"silicon site", ordinarily may not have a phosphorous 5+ cation as a nearest neighbor. The four nearest neighbor lattice sites may therefore be occupied by one silicon and three aluminum cations, two silicon and two aluminum cations, three silicon and one aluminum cations, four silicon cations, or four aluminum cations. As discussed, conventional AEL-type SAPOs with increased silicon concentration, above about 4 molar %, contain undesirable silicon islands, i.e., silicon atoms in the framework having four silicon atom nearest neighbors. The silicon atoms in the preferred composition are physically distributed so that the silicon island concentration is greatly reduced compared with conventional SAPO-11 (i.e., SAPO-11 prepared in accordance with the methods disclosed in the prior art) having the same total silicon concentration.

In a particularly preferred embodiment, the total molar fraction of silicon in the ECR-42 framework is greater than about 0.05, and the number of Si atoms having no Si nearest neighbor ranges from about 0 mol. % to about 100 mol. %, the number of Si atoms having one Si nearest neighbor ranges from about 5 mol. % to about 25 mol. %, the number of Si atoms having two Si nearest neighbors ranges from about 0 mol. % to about 35 mol. %, the number of Si atoms having three Si nearest neighbors ranges from about 0 mol. % to about 25 mol. %, and the number of Si atoms having four Si nearest neighbors ranging from about 0 mol. % to about 25 mol. %, the mol. % being based on the total silica in the ECR-42.

Preferably, the ECR-42 compositions have a Si content ranging from above about 4 mol. % to about 20 mol. %, more preferably from about 5 mol. % to about 15 mol. %, and still more preferably from about 7 mol. % to about 15 mol. %, the silicon content being based on the total amount of Si present in the framework of the molecular sieve composition. Preferably, more than about 50 molar %, and more preferably more than about 90 molar % of the silicon atoms present in the framework do not have four silicon atoms as nearest neighbors in the framework.

The preferred molecular sieve compositions have both a desirably high silicon concentration and a desirable silicon atom dispersion (i.e., a smaller number of silicon islands than would be present in conventional SAPO-11 with the same silicon content) may be formed in accordance with conventional molecular sieve synthesis techniques. The preferred ECR-42 synthesis processes commence with the formation of a silicoaluminophosphate gel having the formula



wherein SURF is a surfactant capable of forming a microemulsion between synthesis solution's organic and aqueous phases. While not wishing to be bound by any theory or model, it is believed that the surfactant helps to solubilize the organic tetraalkylorthosilicate silica source and also increases the available interface between the organic species and inorganic species in the synthesis gel resulting in an improvement in the final silica distribution in the silicoaluminophosphate product. Non-limiting examples of useful surfactants include one or more of long chain alkylamines such as hexadecylamine, tetradecylamine, dodecylamine, decylamine, or octylamine, or dimethyl alkylamine compounds such as dimethylhexadecylamine or dimethyloctylamine, or trimethylalkylammonium salts such as trimethylhexadecylammounium chloride.

TEMP is a template capable of structure directing to AEL such as di-n-propylamine, diisopropylamine, or diethylamine for forming the preferred ECR-42 silicoaluminophosphate molecular sieve.

SOL is a water-soluble organic co-solvent capable of solubilizing the organic silicon source. While not wishing to be bound, it is believed that solubilizing organic silicon sources such as a tetraalkylorthosilicate silicon source into the aqueous synthesis gel improves the final silicon distribution in

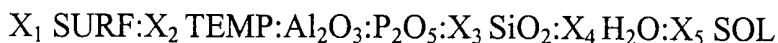
the silicoaluminophosphate product. Non-limiting examples of useful water-soluble organic solvents include one or more of acetone, 1,2-propanediol, 1,3-propanediol, methanol, ethanol, propanol, isopropanol, butanol, or ethylene glycol.

The silicon source material may be any silicon species capable of being dispersed or dissolved in an aqueous synthesis solution. As discussed, where an organic silicon species is employed, a water-soluble organic solvent, SOL, is preferably employed. While not wishing to be bound by any theory or model, it is believed dispersing the silicon species in a low molecular weight form in the silicoaluminophosphate synthesis gel improves silicon distribution of the preferred ECR-42 material formed therefrom. Non-limiting examples of useful silicon source materials include one or more of tetraalkylorthosilicates such as tetramethylorthosilicate, tetraethylorthosilicate, tetrapropylorthosilicate, tetrabutylorthosilicate, and silsesquioxanes having up to twelve Si centers.

X_1 ranges from about 0 to about 0.5, X_2 ranges from about 0.1 to about 4, X_3 ranges from about 0.01 to about 2, X_4 ranges from about 10 to about 100, and X_5 ranges from about 0 to about 30. The silicoaluminophosphate gel may be formed in accordance with the following processes.

(I) Processes using surfactant and water-soluble co-solvent.

Silicoaluminophosphate gels having the formula



may be prepared by combining a phosphorus source such as H_3PO_4 , ammonium phosphate, and mixtures thereof; water; and an aluminum source such as hydrated alumina, organo alumina, pseudo-boehmite, aluminum hydroxide, colloidal alumina, and mixtures thereof, and then adding the template in order to form a homogeneous alumino phosphate mixture. The $\text{P}_2\text{O}_5:\text{Al}_2\text{O}_3$ molar ratio in

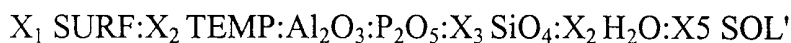
the alumino phosphate mixture preferably ranges from about 0.7 to about 1.3, and more preferably ranges from about 0.9 to about 1.1. The TEMP:Al₂O₃ molar ratio in the alumino phosphate mixture preferably ranges from about 0.1 to about 5, and more preferably from about 0.5 to about 3. A surfactant solution of the surfactant, the -water-soluble organic solvent, and water may then be added to the alumino phosphate mixture along with the organic silicon source and water (if necessary to obtain the desired value of X₄) in order to form a synthesis solution having a SURF:Al₂O₃ molar ratio preferably ranging from about 0 to about 0.5, and more preferably ranging from about 0.05 to about 0.3. The relative amounts of the ingredients for the synthesis solution may be calculated from the desired values of X₁ through X₅. The SOL:Al₂O₃ molar ratio in the synthesis solution preferably ranging from about 0 to about 30, and more preferably ranging from about 4 to about 20, and the SiO₂:Al₂O₃ molar ratio preferably ranging from about 0.01 to about 2.0, and more preferably ranging from about 0.1 to about 0.8. The synthesis solution may then be subjected to gel formation conditions in order to provide the silicoaluminophosphate gel. The ingredients of the synthesis solution, surfactant solution, and alumino phosphate mixture may be mixed in any order. For example, in an alternative embodiment the organic silicon source may be added to a mixture of phosphoric acid and the water-soluble organic solvent. The template and the surfactant may then be added, with the water and aluminum source being added last.

Preferred silicoaluminophosphate crystallization conditions for forming the ECR-42 molecular sieve from the gel include heating the gel in a microwave autoclave for a time sufficient to crystallize the gel. Lower temperature and shorter crystallization times are preferred because such conditions may prevent the formation of undesirable products. Accordingly, the preferred heating times range from about 1 minute to about 5 days, at a temperature ranging from about 100°C to about 250°C, and at a pressure ranging from about 0 bar to about 70 bar. In cases where other products, unreacted gel,

or a mixture thereof is present at the conclusion of the reaction, the molecular sieve may be recovered by a separation process such as centrifugation. The process may also include conventional product washing and drying such as an ethanol rinse, followed by a water rinse, followed by air oven drying at a temperature ranging from about ambient temperature to about 200°C. It should be noted that conventional heating, in for example an air oven or an externally heated autoclave, may be substituted for microwave heating in this process, and that a substantially pure ECR-42 molecular sieve composition having the AEL-topology will result with either heating method. When conventional heating is used, the temperature preferably ranges from about 100°C to about 250°C, and more preferably from 150°C to about 200°C.

(II) Processes using surfactant, hexanol co-solvent and microwave heating.

Silicoaluminophosphate gels having the formula

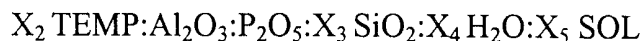


wherein SOL' is a relatively water insoluble organic solvent including solvent mixtures capable forming a microemulsion with water and a surfactant, and including mixtures thereof, may be prepared by according to the method (I) above, except that the water insoluble organic solvent SOL' is substituted for the water-soluble solvent SOL. But for the solvent substitution, the molar ratios of the ingredients in the alumino phosphate mixture and the synthesis solution are as set forth in method (I), and as in that method, the mixing order is not critical.

The preferred ECR-42 molecular sieve composition may be formed from the gel in accordance with the steps set forth in process (I). It should be noted that conventional heating may result in the presence of SAPO-41 material in the molecular sieve composition.

(III) Processes using ethanol with water-soluble cosolvent and no added surfactant.

Silicoaluminophosphate gels of stoichiometry:

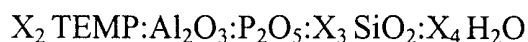


may be prepared by a method identical to method (I), above, except that no surfactant solution is used. An aqueous solution of water, the water-soluble organic solvent, and the silicon source is used instead of the surfactant solution. But for the absence of the surfactant, the molar ratios of the ingredients in the alumino phosphate mixture and the synthesis solution are as set forth in method (I), and as in that method, the mixing order of the ingredients is not critical.

The preferred ECR-42 molecular sieve composition may be formed from the gel in accordance with the steps set forth in process (I).

(IV) Processes using an organic silicon source and no added surfactant or water-soluble organic cosolvent.

Silicoaluminophosphate gels of stoichiometry:



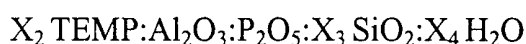
may be prepared by a method identical to method (I), above, except that no surfactant and no co-solvent are employed. An aqueous dispersion of water, and an organic silicon source such as TEOS, including mixtures thereof, is used instead of the surfactant solution.

But for the absence of the surfactant, the molar ratios of the ingredients in the alumino phosphate mixture and the synthesis solution are as set forth in method (I), and as in that method, the mixing order of the ingredients is not critical.

The preferred ECR-42 molecular sieve composition may be formed from the gel in accordance with the steps set forth in process (I). The product may contain SAPO-31.

(V) **Processes for forming SAPO-11 using colloidal silica with no added surfactant and no added solvent.**

Silicoaluminophosphate gels of stoichiometry:



may be prepared by a method identical to method (I), above, except that no surfactant solution is used. An aqueous dispersion of water and the colloidal or fumed silica is used instead of the organic silicon source and surfactant solution. But for the absence of the surfactant and co-solvent, the molar ratios of the ingredients in the alumino phosphate mixture and the synthesis solution are as set forth in method (I), and as in that method, the mixing order of the ingredients is not critical.

A SAPO-11 molecular sieve composition may be formed from the gel in accordance with the steps set forth in process (I). The product may contain SAPO-31.

The preferred ECR-42 molecular sieve compositions are crystalline, and may be prepared with crystal sizes ranging from about 0.001 micrometers to about 1.0 micrometers. Preferably, crystal size ranges from about 0.01 micrometers to about 0.5 micrometers. The synthesis may be adjusted to prepare crystals of decreased size by diluting the synthesis solution with surfactant, co-solvent, and mixtures thereof. Small crystals may also be prepared by conventional methods such as seeding and high shear agitation rates.

Molecular sieve materials prepared in accordance with these methods are useful as catalytic materials. A non-limiting description of the catalytic nature and some catalytic uses of such materials is as follows.

As is known to those skilled in the art, molecular sieve materials may possess an intrinsic or added catalytic functionality, and such materials are frequently referred to as "molecular sieve catalysts". Additional catalytic functionalities may be provided for molecular sieve materials by conventional methods. Such methods are applicable to the preferred molecular sieves, and may be summarized as follows.

Accordingly, the molecular sieve material formed from the gel as set forth above may be calcined to remove the template. The sample may then be allowed to cool, preferably in the presence of a desiccator, and then loaded with a catalytically active species such as Pt by conventional techniques such as incipient wetness.

The preferred catalytic molecular sieve compositions are useful for at least hydrocarbon isomerization, hydrotreating, and cracking such as hydrocracking and fluid catalytic cracking, especially naphtha cracking. The preferred ECR-42 molecular sieve compositions have a dramatically increased catalytic activity for such processes over AEL type SAPO molecular sieves prepared by conventional techniques.

More specifically, the preferred ECR-42 molecular sieves may be used in applications including, but not limited to, catalytic dewaxing, isodewaxing/isomerization, hydrocracking, alkylation of aromatic hydrocarbons (e.g., benzene) with long chain olefins (e.g., C₁₄ olefin), alkylation of aromatic hydrocarbons (e.g., benzene and alkylbenzenes) in presence of an alkylating agent (e.g., olefins, formaldehyde, alkyl halides and alcohols having 1 to about 20 carbon atoms), alkylation of aromatic hydrocarbons (e.g., benzene) with light

olefins to produce short chain aromatic compounds (e.g., alkylation of benzene with propylene to give cumene), transalkylation of aromatic hydrocarbons in the presence of polyalkylaromatic hydrocarbons, isomerization of aromatic feedstock components (e.g., xylene), naphtha cracking to make olefins, oligomerization of straight and branched chain olefins having from about 2 to 5 carbons atoms, disproportionation of aromatics (e.g., the disproportionation of toluene to make benzene and paraxylene), and conversion of naphtha (e.g., C6 - C10) and similar mixtures to highly aromatic mixtures.

In summary, it has been discovered that when AEL-type molecular sieves are preferably synthesized in accord with the methods described herein, that the distribution of Si and therefore the total number and strength of acid sites is quite different, and much higher, than those of conventional SAPO-11. It has also been discovered that the preferred molecular sieves are extremely active and selective for carrying out reactions of hydrocarbon such as alkylation, disproportionation, oligomerization, naphtha cracking and other reactions known to involve molecular sieve catalysts, especially the catalytic dewaxing and isodewaxing of hydrocarbon feeds containing paraffins.

The invention is further exemplified by the following non-limiting examples.

EXAMPLES

I. Preparation of the Preferred ECR-42 Molecular Sieves

Example 1. Preparation of ECR-42 using surfactant, ethanol co-solvent, and microwave heating.

A silicoaluminophosphate gel of stoichiometry:

.16 CA: DPA: Al₂O₃: P₂O₅: 0.2 SiO₂: 50.3 H₂O: 7.3 ethanol

(where CA is hexadecylamine) was prepared by mechanically mixing together 28.6 grams of H_3PO_4 (85%), 25 grams of water and 17 grams of Catapal B alumina (Vista Chemical Co., Houston, Texas 74% Al_2O_3 , 26% H_2O) for 45 min. in a 500 ml Teflon bottle until homogenous. To this mixture was added 12.7 grams of DPA, and the mixture was then stirred for 30 minutes. A surfactant solution was prepared by mixing 2.4 grams of hexadecylamine in 20.8 grams of ethanol and 12.5 grams of water. This solution was added to half of the di-n-propylamine/aluminophosphate mixture along with 2.6 grams of TEOS and 11.2 grams of water. The resulting mixture was stirred for about 3 min. in a 125 ml blender. This gel was divided between three CEM XP-1500 Teflon microwave autoclaves (about 28.5 grams each), available from CEM Corp., Matthews, NC. The three autoclaves were heated for 5 min. in a microwave oven to a pressure of 485 psi. After 5, 30 and 60 min. at 485 psi the samples were removed from the oven. The ECR-42 products were recovered by centrifugation and washed twice with ethanol and then twice with de-ionized water. The resulting products were dried in an air oven at 115°C and the powder X-ray diffraction pattern measured which showed the products to be pure SAPO-11, see figure 4(a). Elemental analysis of the 60 min. sample gave: 17.3% Al; 17.0% P; 1.68% Si, representing a product stoichiometry of $\text{Si}_{0.05}\text{Al}_{.51}\text{P}_{.44}$.

Example 2. Preparation of ECR-42 using surfactant, hexanol co-solvent, and microwave heating.

A silicoaluminophosphate gel of stoichiometry:

.16 CA: DPA: Al_2O_3 : P_2O_5 : 0.4 SiO_2 : 39 H_2O : 4.5 hexanol

(where CA is hexadecylamine) was prepared by mechanically mixing together 28.6 grams of H_3PO_4 (85%), 25 grams of water and 17 grams of Catapal B alumina (Vista, 74% Al_2O_3 , 26% H_2O) for 2 hr. in a 500 ml Teflon bottle until homogenous. To this mixture was added 12.7 grams of DPA and the mixture

was then stirred for 30 minutes. A surfactant solution was prepared by mixing 4.8 grams of hexadecylamine in 56.9 grams of hexanol and 25 grams of water. This solution was added to the di-n-propylamine/aluminumphosphate mixture along with 10.3 grams of TEOS and 22.3 grams of water. The resulting mixture was stirred for about 3 min. in a 125 ml blender. This gel was divided between two 125 ml Teflon-lined autoclaves and a CEM XP-1500 Teflon microwave autoclave (about 28.5 grams). The two autoclaves were heated at 195°C for 24 and 40 hrs. in a conventional oven. The microwave autoclave was heated for 5 min. in a microwave oven to a pressure of 485 psi. After 60 min. at 485 psi the sample was removed from the oven. The products were recovered by centrifugation and washed twice with ethanol and then twice with de-ionized water. The resulting ECR-42 products were dried in an air oven at 115°C and the powder X-ray diffraction pattern measured showed that the product of the microwave heating was pure SAPO-11 (figure 4(c)) and the product of the conventionally heated samples to be SAPO-11 and SAPO-41 (figure 4(b)). Elemental analysis of the microwave heated sample gave: 14.5% Al; 16.4% P; 2.84% Si, representing a product stoichiometry of $\text{Si}_{0.09}\text{Al}_{.46}\text{P}_{.45}$ and the conventionally heated sample (40 hrs.) gave 15.4% Al; 16.4% P; 3.36% Si, representing a product stoichiometry of $\text{Si}_{1.10}\text{Al}_{.47}\text{P}_{.43}$.

Example 3. Preparation of ECR-42 using surfactant, ethanol co-solvent, and microwave heating.

A silicoaluminumphosphate gel of stoichiometry:

.16 CA: DPA: Al_2O_3 : P_2O_5 : 0.2 SiO_2 : 39.1 H_2O : 7.3 ethanol

(where CA is hexadecylamine) was prepared by mechanically mixing together 71.5 grams of H_3PO_4 (85%), 62.5 grams of water and 42.4 grams of Catapal B alumina (Vista, 74% Al_2O_3 , 26% H_2O) for 45 min. in a 1000 ml Teflon bottle until homogenous. To this mixture was added 31.7 grams of DPA and the

mixture was then stirred for 30 minutes. A surfactant solution was prepared by mixing 12 grams of hexadecylamine in 104 grams of ethanol and 62.5 grams of water. This solution was added to the di-n-propylamine/aluminophosphate mixture along with 13 grams of TEOS and 56 grams of water. The resulting mixture was stirred for about 3 min. in a 250 ml blender. One half of this gel was divided between five CEM XP-1500 Teflon microwave autoclaves (about 42.5 grams each). The five autoclaves were heated for 10 min. in a microwave oven to a pressure of 485 psi. After 30 min. at 485 psi the samples were removed from the oven. The ECR-42 products were combined and then recovered by centrifugation and washed twice with ethanol and then twice with de-ionized water. The resulting product was dried in an air oven at 115°C and the powder X-ray diffraction pattern measured which showed the product to be pure SAPO-11. See figure 4(d). Elemental analysis gave: 15.96% Al; 17.3% P; 1.638% Si, representing a product stoichiometry of $\text{Si}_{1.05}\text{Al}_{4.9}\text{P}_{4.6}$.

Example 4. Preparation of ECR-42 using surfactant, ethanol co-solvent and microwave heating.

A silicoaluminophosphate gel of stoichiometry:

.16 CA: DPA: Al_2O_3 : P_2O_5 : 0.4 SiO_2 : 39.1 H_2O : 7.3 ethanol

(where CA is hexadecylamine) was prepared by adding 6.5 grams more of TEOS to the second half of the gel prepared in Example 3. After thoroughly homogenizing in a blender this gel was divided between five CEM XP-1500 Teflon microwave autoclaves (about 42.5 grams each. The five autoclaves were heated for 10 min. in a microwave oven to a pressure of 485 psi. After 30 min. at 485 psi the samples were removed from the oven. The ECR-42 products were combined and then recovered by centrifugation and washed twice with ethanol and then twice with de-ionized water. The resulting product was dried in an air oven at 115°C and the powder X-ray diffraction pattern measured which showed

the product to be pure SAPO-11. See figure 4(e). Elemental analysis gave: 15.23% Al; 17.2% P; 2.795% Si, representing a product stoichiometry of $\text{Si}_{.08}\text{Al}_{.46}\text{P}_{.46}$.

Example 5 Preparation of ECR-42 using surfactant, ethanol co-solvent, and conventional heating.

A silicoaluminophosphate gel of stoichiometry:

.16 CA: 2 DPA: Al_2O_3 : P_2O_5 : 0.4 SiO_2 : 40 H_2O : 7.5 ethanol

(where CA is hexadecylamine) was prepared by mixing together 25.4 grams of H_3PO_4 (85%), 35 grams of water, and 15.2 grams of Catapal A alumina (Vista, 74% Al_2O_3 , 26% H_2O) for 2 min. in a 250 ml plastic beaker. To this mixture was added 22.3 grams of DPA and the mixture was then stirred for 3 minutes. A surfactant solution was prepared by mixing 4.25 grams of hexadecylamine in 38.1 grams of ethanol for thirty minutes. This solution was added to the di-n-propylamine/aluminophosphate mixture along with 9.18 grams of TEOS and 30.6 grams of water. The resulting mixture was homogenized for about 5 min. in a 250 ml blender. About 80 g of this gel was placed in a 125 Teflon lined autoclave. The autoclave was heated for 24 hours in an air oven at 195°C. After cooling, the ECR-42 product was recovered by centrifugation and washed twice with ethanol and then twice with de-ionized water. The resulting product was dried in an air oven at 115°C and the powder X-ray diffraction pattern measured which showed the product to be pure SAPO-11, see figure 5(a). Elemental analysis gave: 17.2% Al; 17.1% P; 3.052% Si, representing a product stoichiometry of $\text{Si}_{.084}\text{Al}_{.491}\text{P}_{.425}$.

Example 6. Preparation of ECR-42 using TEOS, ethanol co-solvent and conventional heating.

A silicoaluminophosphate gel of stoichiometry:

2 DPA: Al₂O₃: P₂O₅: 0.4 SiO₂: 40 H₂O: 7.5 ethanol

was prepared by mixing together 28.9 grams of H₃PO₄(85%), 40 grams of water and 17.3 grams of Catapal A alumina (Vista, 74%Al₂O₃, 26% H₂O) for 2 min. in a 250 ml plastic beaker. To this mixture was added 25.4 grams of DPA and the mixture stirred for 3 minutes. To the di-n-propylamine/aluminophosphate mixture was added 43.3 grams of ethanol with 10.4 grams of TEOS and 34.7 grams of water. The resulting mixture was homogenized for about 5 min. in a 250 ml blender. About 80 g of this gel was placed in a 125 Teflon lined autoclave. The autoclave was heated for 24 hours in an air oven at 195°C. After cooling, the ECR-42 product was recovered by centrifugation and washed once with ethanol and then three times with de-ionized water. The resulting product was dried in an air oven at 115°C and the powder X-ray diffraction pattern measured which showed the product to be pure SAPO-11. See figure 5(b). Elemental analysis gave: 17.94% Al; 20.1% P; 2.56% Si, representing a product stoichiometry of Si_{1.065}Al_{4.73}P_{.462}.

Example 7. Preparation of ECR-42 using TEOS and conventional heating.

A silicoaluminophosphate gel of stoichiometry:

1.5 DPA: Al₂O₃: P₂O₅: 0.4 SiO₂: 50 H₂O

was prepared by mixing together 25.1 grams of H₃PO₄(85%), 36 grams of water and 15 grams of Catapal A alumina (Vista, 74%Al₂O₃, 26% H₂O) for 2 min. in a 250 ml plastic beaker. To this mixture was added 16.5 grams of DPA, and the mixture was stirred for 3 minutes. To the di-n-propylamine/aluminophosphate mixture was added 9.1 grams of TEOS (tetraethylorthosilicate) and 48.4 grams of water. The resulting mixture was homogenized for about 5 min. in a 250 ml blender. One half of this gel was placed in a 125 Teflon-lined autoclave. The autoclave was heated for 24 hours in an air oven at 220°C. After cooling, the ECR-42 product was recovered by centrifugation and washed four times with de-

ionized water. The resulting product was dried in an air oven at 115°C and the powder X-ray diffraction pattern measured which showed the product to be SAPO-11 and SAPO-41. See figure 5(c). Elemental analysis gave: 18.8% Al; 19.7% P; 2.70% Si, representing a product stoichiometry of $\text{Si}_{0.067}\text{Al}_{0.49}\text{P}_{0.445}$.

Example 8. Preparation of SAPO-11 using colloidal silica.

A silicoaluminophosphate gel of stoichiometry:

1.5 DPA: Al_2O_3 : P_2O_5 : 0.4 SiO_2 : 50 H_2O

was prepared by mixing together 31.4 grams of H_3PO_4 (85%), 80 grams of water and 18.8 grams of Catapal B alumina (Vista, 74% Al_2O_3 , 26% H_2O) and mechanically stirring for 0.5 hour in a 500 ml Teflon bottle. To this mixture was added 20.7 grams of DPA (di-n-propylamine) and 8.19 g solution of aqueous colloidal silica (duPont Ludox AS-40, 40% SiO_2) and stirred for 1.5 hours. Finally, 20.9 grams of water were added, and the resulting mixture was homogenized for about 5 min. in a 250 ml blender. This gel was divided between two 125 Teflon lined autoclaves. The autoclaves were heated for 24 hours in an air oven at 195°C. After cooling, the products were combined and then recovered by centrifugation and washed four times with de-ionized water. The resulting product was dried in an air oven at 115°C and the powder X-ray diffraction pattern measured which showed the product to be SAPO-11 and SAPO-31. See figure 5(d). Elemental analysis gave a product stoichiometry of $\text{Si}_{0.056}\text{Al}_{0.502}\text{P}_{0.442}$.

Examples 9. Preparation of ECR-42 using various alcohol co-solvents and tetraalkylorthosilicates.

A series of silicoaluminophosphate gels of stoichiometry:

X_1 CA: 2 DPA: Al_2O_3 : P_2O_5 : 0.4 SiO_2 : 40 H_2O : X_5 SOL

where CA is hexadecylamine and DPA is di-n-propylamine) were prepared by mixing together the appropriate amounts of H_3PO_4 (85%), water and Catapal A alumina (Vista, 74% Al_2O_3 , 26% H_2O) for 2 min. in a 250 ml plastic beaker. To this mixture was added the appropriate amount of DPA (di-n-propylamine) and the mixture stirred for 3 minutes. Then the appropriate amount of alcohol, tetraalkylorthosilicate and water were added to obtain the stoichiometry as indicated in Table 1. The resulting mixtures were homogenized for about 5 min. in a 250 ml blender. One half of this mixture was transferred to 125 Teflon-lined autoclave. To the other half was mixed in 1.5 grams of hexadecylamine and also transferred to 125 Teflon-lined autoclave. The autoclaves were heated for 24 hours in an air oven at 195°C. After cooling the product was recovered by centrifugation and washed twice with ethanol and then twice with de-ionized water. The resulting products were dried in an air oven at 115°C and the powder X-ray diffraction patterns measured which showed the products to be pure ECR-42. Elemental analyses and product stoichiometries are listed in Table 2.

TABLE 1

Experiment Number	X_1	X_5	SOL	Si Source
9	0	11.5	Methanol	tetraethylorthosilicate
10	.16	11.5	Methanol	tetraethylorthosilicate
11	0	6.1	n-propanol	tetraethylorthosilicate
12	.16	6.1	n-propanol	tetraethylorthosilicate
13	0	6.1	i-propanol	tetraethylorthosilicate
14	.16	6.1	i-propanol	tetraethylorthosilicate
15	0	5.0	n-butanol	tetraethylorthosilicate
16	.16	5.0	n-butanol	tetraethylorthosilicate
17	0	8	ethanol	tetramethylorthosilicate
18	.16	8	ethanol	tetramethylorthosilicate
19	0	8	ethanol	tetrapropylorthosilicate
20	.16	8	ethanol	tetrapropylorthosilicate

TABLE 2

Experiment Number	%Si	%Al	%P	Molar Si	Molar Al	Molar P
9	3.213	18.40	19.37	0.080	0.480	0.440
10	3.870	17.63	13.68	0.112	0.530	0.358
11	2.813	20.73	22.22	0.063	0.484	0.452
12	3.542	18.27	15.00	0.098	0.526	0.376
13	3.147	18.09	19.42	0.080	0.476	0.445
14	3.550	18.65	15.23	0.097	0.528	0.376
15	2.668	18.60	19.70	0.067	0.485	0.448
16	3.360	19.38	15.78	0.089	0.533	0.378
17	3.148	18.31	19.35	0.079	0.479	0.441
18	4.194	18.35	12.98	0.120	0.545	0.336
19	3.124	18.00	19.24	0.079	0.477	0.444
20	3.730	19.52	15.90	0.097	0.528	0.375

Example 10.

The materials of experiments 1-20 were subject to ^{29}Si MAS NMR to determine the relative amounts of silicon in the five different environments, i.e., Si(4Si), Si(3Si), Si(2Si), Si(1Si), Si(0Si). The ^{29}Si Magic-angle spinning (MAS) nuclear magnetic resonance (NMR) spectra with proton decoupling were recorded on a Chemagnetics CMXII 500 spectrometer at 99.3 MHz using a 7.5 mm probe spinning at 6 kHz. Typical run conditions were: 800 to 10000 acquisitions; 4 or 5 sec. pulse width; 60 to 120 sec. relaxation delay. Chemical shifts were measured relative to tetramethylsilane (TMS). The spectra were then deconvoluted into the five silicon environments. The molar fractions of each silicon environment are given in Table 3.

TABLE 3

Example	4Al,0Si	3Al,1Si	2Al,2Si	1Al,3Si	0Al,4Si
1	54	21	13	9	4
2					
3					
4					
5	49	19	16	11	5
6	33	22	22	13	10
7					
8	36	17	7	17	23
9	24	31	18	18	9
10	24	44	15	14	3
11					
12					
13	32	32	22	5	9
14	41	31	6	12	10
15					
16					
17	35	15	10	18	22
18					
19					
20					

II. Discussion

A. Crystallite Size and Shape

It is well known that zeolite morphology has a major influence on the reactivities of zeolites. Such effects become more important with smaller pores and larger molecules. At very small crystal sizes, the exterior surface may play a major role in processes using such zeolites, especially the location of exterior pore openings with respect to the catalytically active site(s).

Powder X-ray diffraction may be used for determining crystallite size (and often shape) of materials less than about 1000Å by the relative broadening of peaks in the X-ray diffraction pattern. When the shape of the crystallites is anisotropic (different dimensions in different directions), then the resulting powder diffraction pattern may have both broad and sharp peaks if one

or two of the dimensions of the crystals are small enough to show peak broadening. Those reflections coming from Miller planes in the crystals corresponding to the short dimension will have broader peaks than those coming from Miller planes corresponding to the longer dimension will have sharper peaks.

SAPO-11 has an orthorhombic crystal structure, a structure which is frequently associated with spatial anisotropies. The dimensions of the unit cell are $a = 8.4\text{\AA}$, $b = 18.5\text{\AA}$ and $c = 13.5\text{\AA}$, and the structure is such that the 10-ring channels run along the a dimension [100]. If the shape of the crystallites was platelet in shape such that the a dimension was significantly smaller than the b and c dimensions, one would expect the X-ray diffraction pattern to show the peaks corresponding to the $(h00)$ reflections to be broader than those of the $(0kl)$ reflections. This would result in those peaks corresponding to the $(h00)$ reflections to have shorter peak height relative to the $(0kl)$ peaks, not due to a reduction on peak intensity, but to an increase in the peak width.

While not wishing to be bound by any model or theory, it is believed that an increased catalytic activity for a molecular sieve catalyst would result when a large number of channel openings are available to allow for efficient transportation of the reacting species and products to and from the active sites inside of, or at the pore mouth of, the molecular sieve crystal. The preferred ECR-42 molecular sieve catalysts have this desirable spatial anisotropy (preferred orientation), as shown in the powder x-ray diffraction results in Figures 4 and 5. The (200) reflections are clearly broader in width and shorter in height than surrounding $(0kl)$ reflections as the reactants are changed from TEOS (Example 7) to TEOS and ethanol co-solvent (Example 6) to TEOS, surfactant, and ethanol co-solvent (Example 5). The SEM micrograph of Figure 6(a) clearly shows the preferred thin platelets present in the product of Example 4. The platelets have a thickness less than 50 nm, well within the threshold for

seeing X-ray diffraction peak broadening, and consistent with the diffraction patterns in Figures 4 and 5. Figure 6(b) on the other hand, shows the undesirable thick platelet morphology that results in conventionally prepared SAPO-11.

B. Silicon Distribution by ^{29}Si NMR

As discussed, while the preferred molecular sieve materials have an X-ray diffraction pattern with line positions corresponding to the AEL topology, they are physically and chemically distinct from known SAPO-11 materials because of differences in the local atomic arrangement of the component atoms. Conventional powder x-ray diffraction analysis not a sufficient analytical tool to determine the structural differences between SAPO-11 and the preferred ECR-42 molecular sieves partially because of x-ray line broadening, as discussed above. Consequently, in order to determine the above-noted structural differences, it is necessary to use an analytical tool which can discriminate between the atomic environment in the overall crystal structure such as ^{29}Si MAS NMR, which is well known in the art. An example of how the number of Si nearest neighbors may be derived from NMR data is shown in Figure 5 and Table 3.

As discussed, it is believed that is the silicon distribution with the crystal is one important characteristic influencing catalytic activity in SAPO materials. Since silicon is responsible for the acidity in SAPO materials, it is desirable that silicon be well dispersed in the aluminophosphate framework for high catalytic activity. It is known that increasing silicon concentration in conventionally prepared SAPO-11 results in the formation of large silicon islands as shown in Figure 3(b) (low Si concentration, some island formation) and Figure 3(c) (high Si concentration, with extensive island formation). Figure 3(a) shows that still lower Si concentration results in undesirable isolated silicon atoms. Although large silicon islands are undesirable because those silicon atoms at the interior of the islands are catalytically inactive, small silicon-rich

regions present in the preferred ECR-42 material are desirable because the strongest acid sites are believed to form at the borders of the Si -rich regions and the aluminophosphate domains. This is because those silicons at the borders have fewer aluminum atoms as nearest neighbors, which leads to decreasing acidity resulting from aluminum's lower electronegativity.

The distribution of silicon in SAPOs may be measured by ^{29}Si NMR spectroscopy. For example, in SAPO-11, it is known that those silicon atoms having 4 aluminums and 0 silicons (4Al, 0Si) as nearest neighbors show as NMR resonance at -90 to -93 ppm chemical shift relative to tetramethylsilane (TMS). Therefore, the molar percent of Si atoms with zero, one, two, three, and four Si atom nearest neighbors may be obtained, for example, by deconvoluting the integrated intensities from ^{29}Si NMR measurements. The Si (4Al, 0Si) and the other silicon environments are illustrated schematically in Figure 1 and listed in Table 4:

TABLE 4

Si Environment	4Al, 0Si	3Al, 1Si	2Al, 2Si	1Al, 3Si	0Al, 4Si
Chemical shift in ppm from TMS	-86 to -94 ppm	-95 to -99 ppm	-100 to -105 ppm	-106 to -108 ppm	-109 to -115 ppm

It is clear from the above chart that for well dispersed silicon, a ^{29}Si NMR spectra would have high intensity for the (4Al, 0Si) resonance at -86 to -94 ppm, while a poorly dispersed silicon would have high intensity at the (0Al, 4Si) resonance at -109 to -115 ppm. Those silicons located at the border of the silicon atoms having the highest acid strength would show high relative intensity for the (3Al, 1Si), (2Al, 2Si), and (1Al, Si) resonances. Figure 7 shows ^{29}Si NMR spectra of the preferred ECR-42 molecular sieve with framework Si contents of 4, 8, and 13 molar %. The spectra show the excellent silicon dispersion in the preferred ECR-42 material. By way of comparison, Figure 3 shows that Si islanding is avoided in conventional SAPO-11 at very low Si

concentrations only. See Figure 3(b). At Si content above about 0.04 mole fraction, substantial Si islanding occurs. See Figure 3(a) and (c). In particular, Figure 7-B, shows that the preferred ECR-42 molecular sieve does not show undesirable Si island formation at higher Si concentrations than in the conventionally prepared sample shown in Figure 3(a).

It should be noted that due to the characteristic of the experimental NMR spectra, the deconvolutions shown in Figure 2 are not unique, but should be chosen with chemical sense. Therefore, the simulation results constitute an estimation of the different structural units contributing to the spectra, and may be used to illustrate Si atom coordination differences that are evident from the experimental data.

Moreover, the preferred molecular sieve, ECR-42, may contain Si concentrations as low as about 1 wt.%, and it is often difficult to obtain high quality NMR data from samples of such a low concentration over the whole range of chemical shifts. Consequently, some degree of uncertainty may be introduced into the deconvolutions required to determine the number of Si atoms with zero, one, two, three and four Si nearest neighbors. Nevertheless, the intensity of the -109 to -115 ppm region is sufficient in molecular sieve materials of even 1 wt.% to determine the approximate molar percent of Si framework atoms having four Si atom nearest neighbors, i.e., Si atoms located in Si islands.

CLAIMS:

1. A composition comprising a molecular sieve with a framework tetrahedra of silicon, aluminum, and phosphorus, and designated ECR-42, the composition having a total silicon amount ranging from above about 4 molar percent to about 20 molar percent, a total aluminum amount ranging from about 40 molar percent to about 55 molar percent, and a total phosphorus amount ranging from about 30 molar percent to about 50 molar percent, the molar percents being based on the total amount of aluminum, phosphorus, and silicon present in the composition, and the molecular sieve having the topology AEL and being isostructural with conventional SAPO-11, wherein

(a) the silicon present in the molecular sieve and the conventional SAPO-11 is distributed among silicon sites in the framework tetrahedra, each site having a first, a second, a third, and a fourth next nearest neighbor position, and each next nearest neighbor position being independently occupied by one atom selected from silicon, and aluminum, and

(b) the molecular sieve has a substantially smaller number of silicon sites having silicon atoms among all four next nearest neighbor positions than the conventional SAPO-11 having the same total silicon amount.

2. A composition comprising a molecular sieve with a framework tetrahedra of silicon, aluminum, and phosphorus, the composition having a total silicon amount ranging from about 4 molar percent to about 20 molar percent, a total aluminum amount ranging from about 40 molar percent to about 55 molar percent, and a total phosphorus amount ranging from about 30 molar percent to about 50 molar percent, the molar percents being based on the total amount of aluminum, phosphorus, and silicon present in the composition, and the molecular sieve having the topology AEL and being isostructural with conventional SAPO-11, wherein

- 31 -

- (a) the molecular sieve has a first number of Si atoms coordinated as Si(4Si),
 - (b) the conventional SAPO-11 with the Si amount has a second number of Si atoms coordinated as Si(4Si), and
 - (c) the first number of Si atoms is substantially less than the second number of Si atoms.
3. A method for forming a molecular sieve composition, the method comprising:
- (a) combining at least one water-dispersible phosphorus source material, at least one template capable of structure directing to AEL, water, and a water-soluble organic silicon source material under gellation conditions in order to form a gel, and then
 - (b) reacting the gel under molecular sieve synthesis conditions.
4. A molecular sieve product formed by
- (a) combining at least one water-dispersible phosphorus source material, at least one template capable of structure directing to AEL, water, and a water-dispersible silicon source material under gellation conditions in order to form a gel, and then
 - (b) reacting the gel under molecular sieve synthesis conditions.
5. A silicoaluminophosphate gel formed by combining under gellation conditions at least one water-dispersible silicon source material, at least one template capable of structure-directing to AEL, at least one water-dispersible phosphorus source material, and water.

Figure 1

P	Al	P	Al	P	Al	P	Al	P	Al
Al	P	Al	Si	Al	P	Al	P	Al	P
P	Al	P	Al	P	Al	Si	Al	P	Al
Al	P	Al	P	Al	Si	Si	Si	Al	P
P	Al	P	Al	P	Al	Si	Al	P	Al
Al	P	Al	Si	Al	P	Al	P	Al	P
P	Al	Si	Si	Si	Al	P	Al	P	Al
Al	P	Al	Si	Si	Si	Al	P	Al	P
P	Al	Si	Si	Si	Al	P	Al	P	Al
Al	P	Al	Si	Si	Si	Al	P	Al	P
P	Al	Si	Si	Si	Al	P	Al	P	Al
Al	P	Al	Si	Al	P	Al	P	Al	P
P	Al	P	Al	P	Al	P	Al	P	Al
Al	P	Al	P	Al	P	Al	P	Al	P

FIGURE 2

Simulated ^{29}Si NMR for SAPO-11

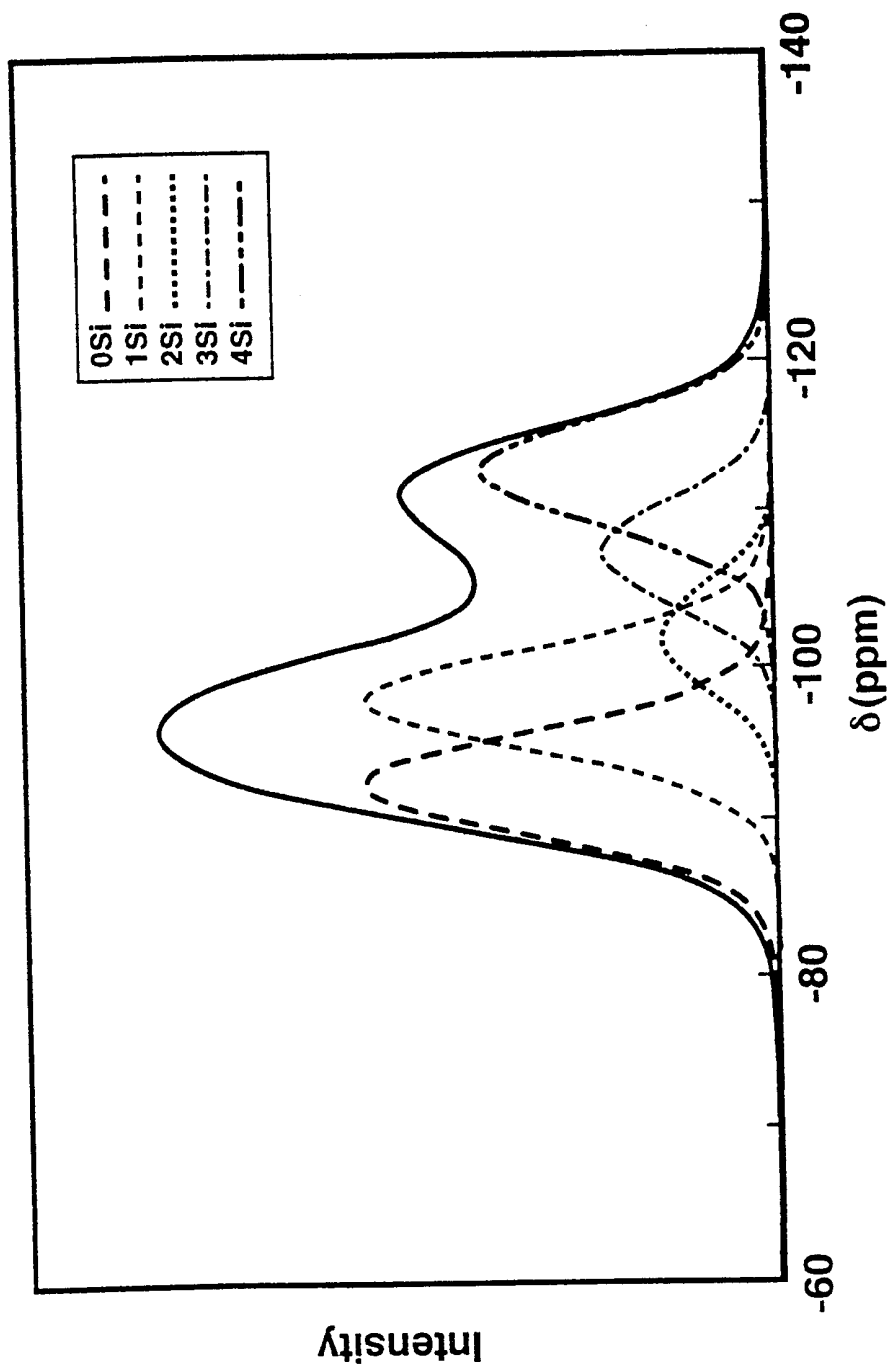


Figure 3

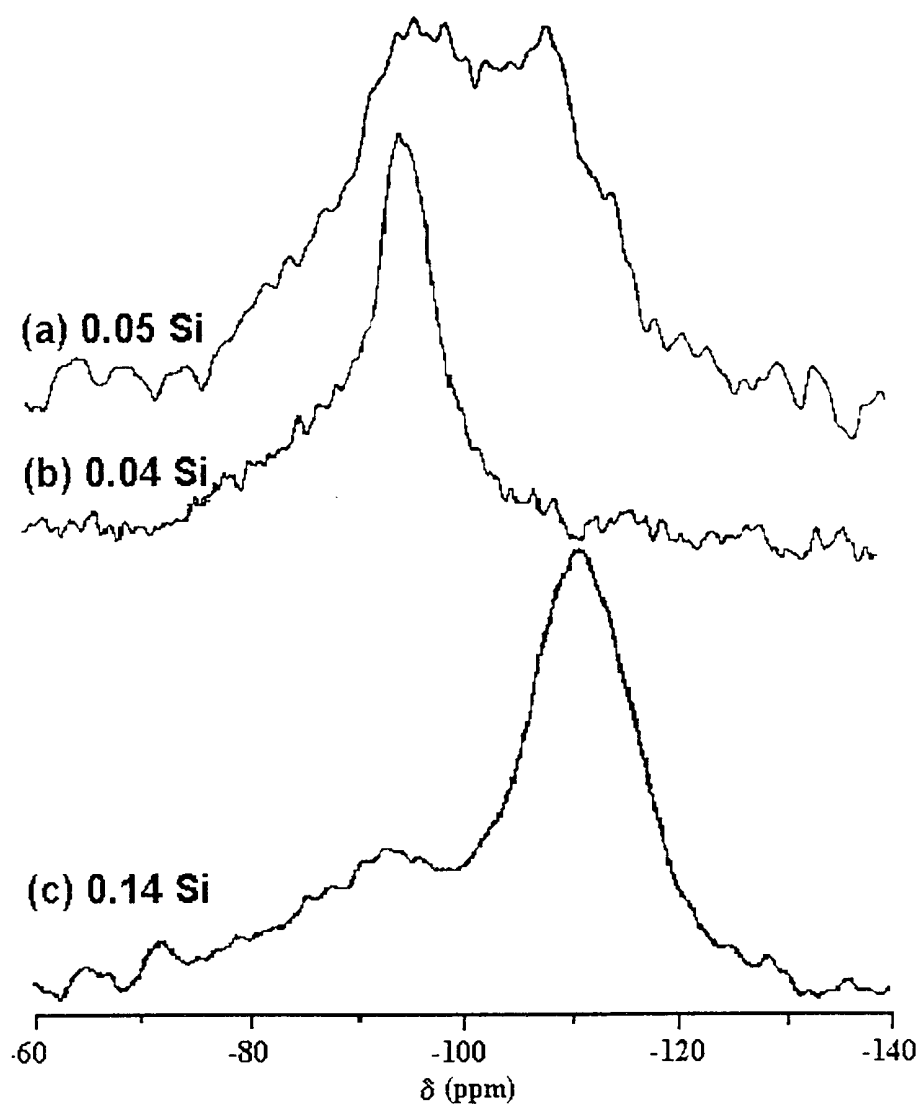


Figure 4

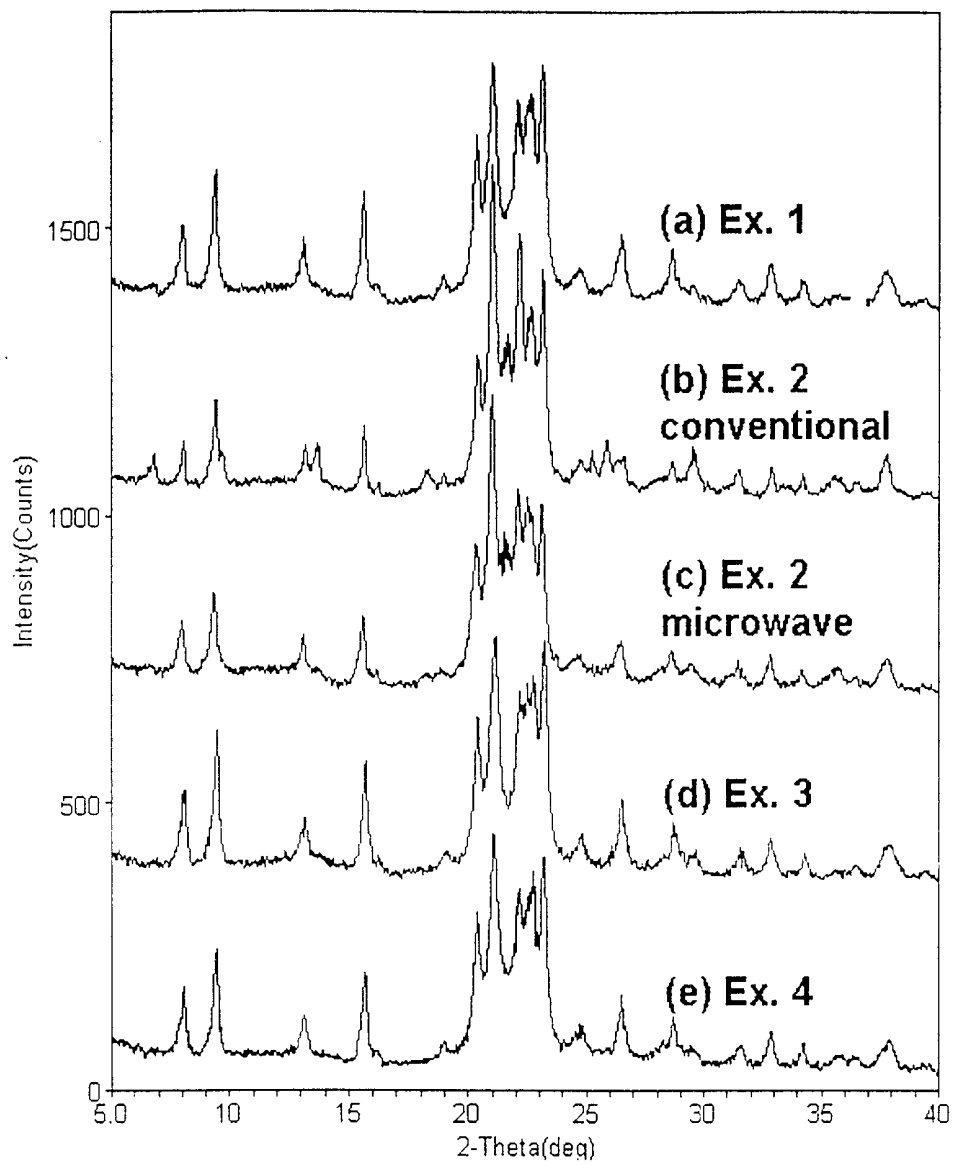


Figure 5

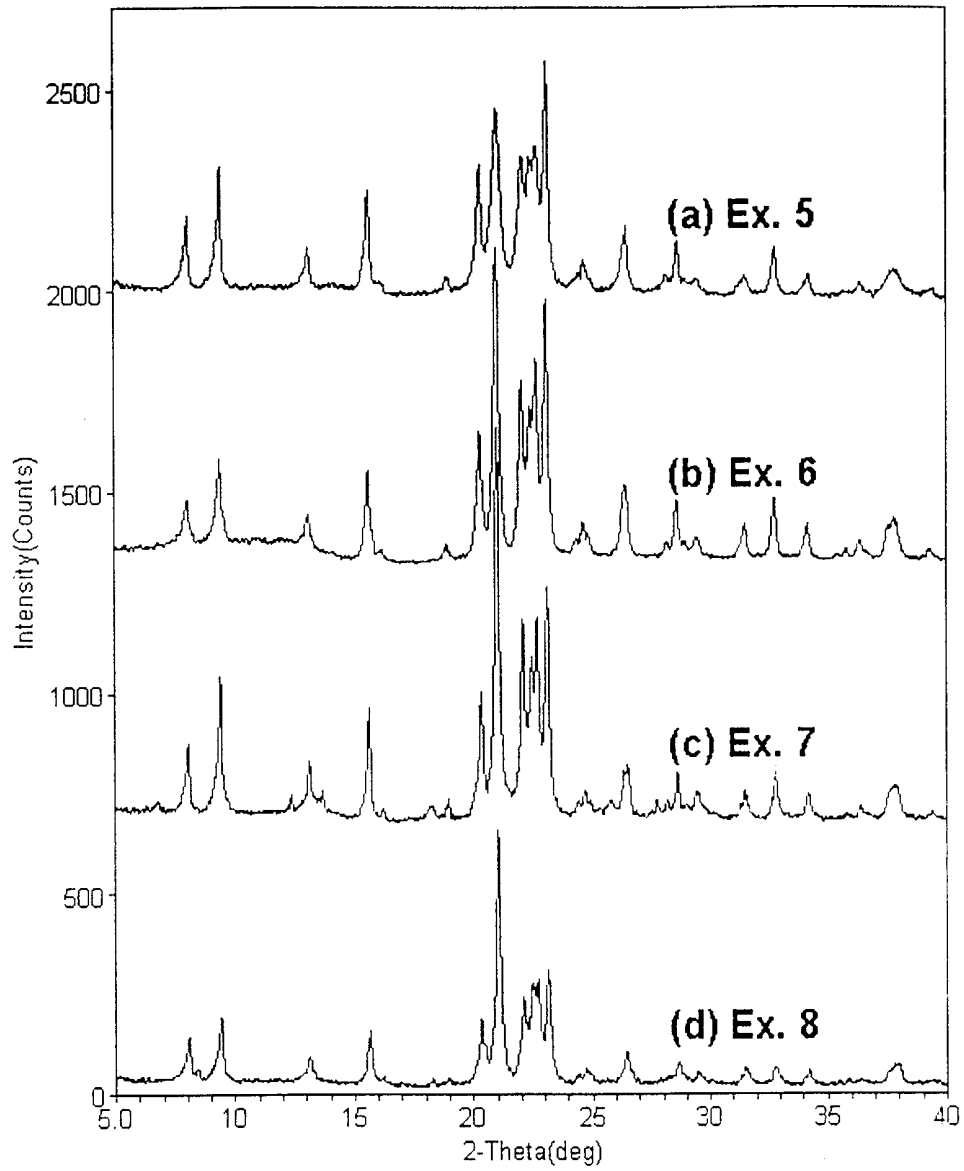


Figure 6

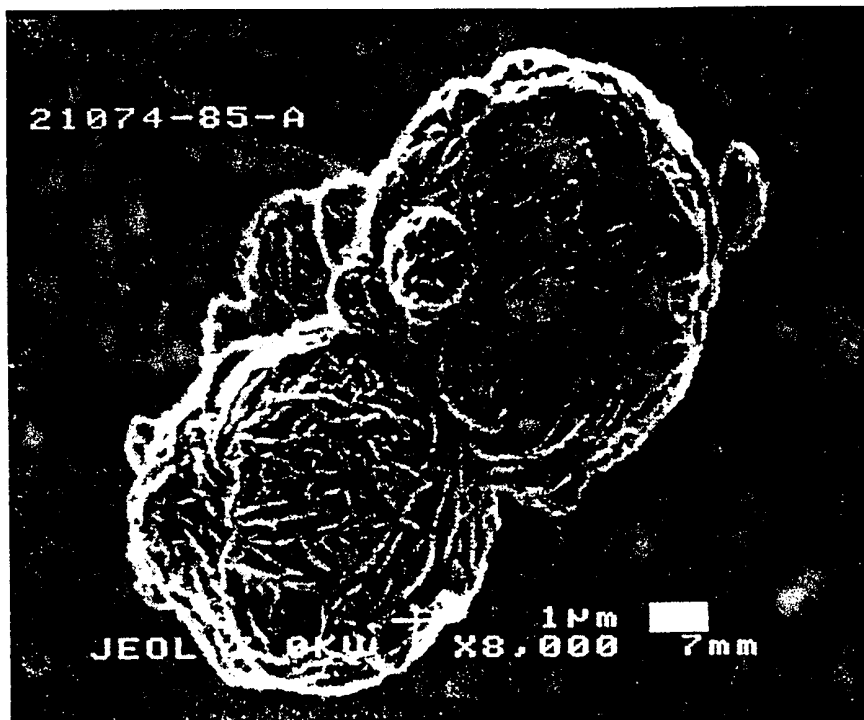
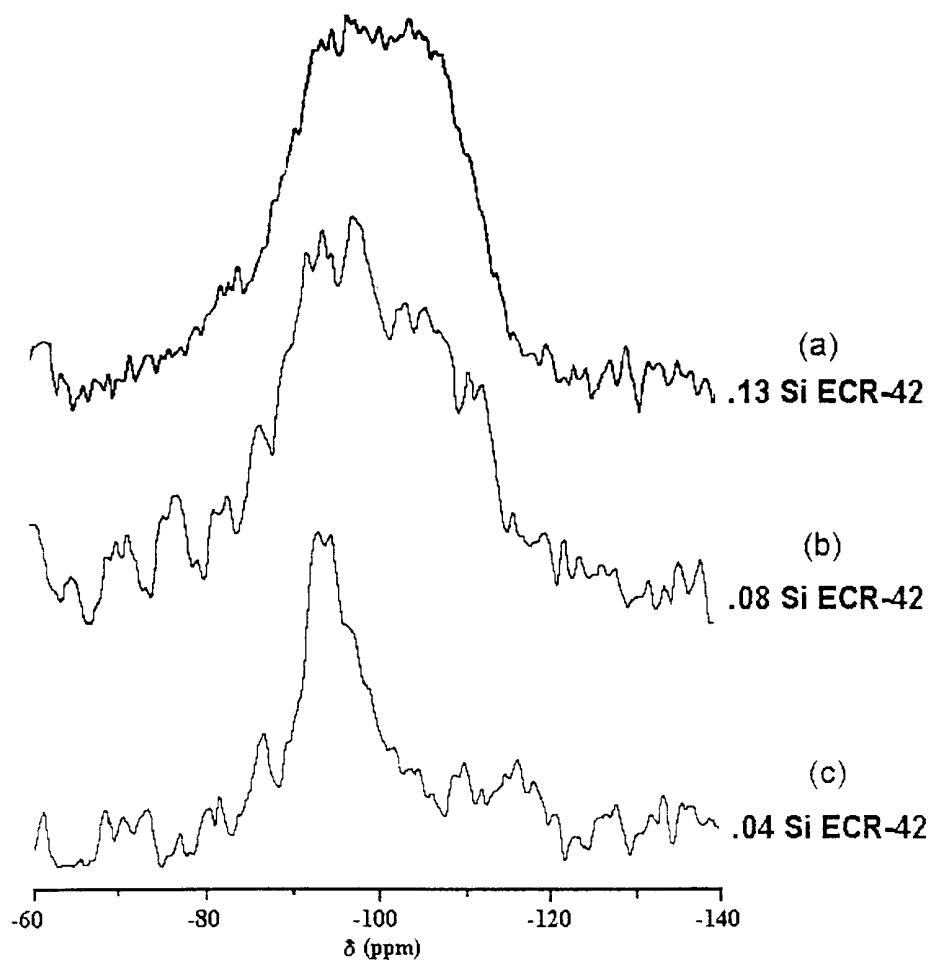


Figure 6(a) – SEM of product form Example 4.



Figure 6(b) – SEM of conventional SAPO-11 material

Figure 7



INTERNATIONAL SEARCH REPORT

International application No.
PCT/US99/11469

A. CLASSIFICATION OF SUBJECT MATTER

IPC(6) :C01B 37/08
US CL :423/306, 702, DIG.30; 502/214
According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 423/306, 702, DIG.30; 502/214

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X --- A	US 5,096,684 A (GUTH et al.) 17 March 1992, col. 2, line 25 to col. 3, line 11, col. 3, lines 52-58, col. 11, lines 32-33.	3-5 --- 1-2
X	US 4,440,871 A (LOK et al.) 03 April 1984, col. 4, lines 17-32, col. 6 lines 30-31, col. 7, line 63 to col. 8, line 9.	3-5
A	US 4,943,424 A (MILLER) 24 July 1990.	1-2

Further documents are listed in the continuation of Box C. See patent family annex.

* Special categories of cited documents:	*T* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
A document defining the general state of the art which is not considered to be of particular relevance	*X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
E earlier document published on or after the international filing date	*Y* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
L document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)	*G* document member of the same patent family
O document referring to an oral disclosure, use, exhibition or other means	
P document published prior to the international filing date but later than the priority date claimed	

Date of the actual completion of the international search 03 AUGUST 1999	Date of mailing of the international search report 17 AUG 1999
Name and mailing address of the ISA/US Commissioner of Patents and Trademarks Box PCT Washington, D.C. 20231 Facsimile No. (703) 305-3230	Authorized officer <i>B. Sample</i> DAVID SAMPLE Telephone No. (703) 308-0661