



- (51) **International Patent Classification:**
C07C 39/04 (2006.01) *C07C 47/546* (2006.01)
- (21) **International Application Number:**
PCT/US2012/031994
- (22) **International Filing Date:**
3 April 2012 (03.04.2012)
- (25) **Filing Language:** English
- (26) **Publication Language:** English
- (30) **Priority Data:**
61/471,450 4 April 2011 (04.04.2011) US
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- (81) **Designated States (unless otherwise indicated, for every kind of national protection available):** AE, AG, AL, AM, AO, AT, AU, AZ, BA, BB, BG, BH, BR, BW, BY, BZ, CA, CH, CL, CN, CO, CR, CU, CZ, DE, DK, DM, DO, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, GT, HN, HR, HU, ID, IL, IN, IS, JP, KE, KG, KM, KN, KP, KR, KZ, LA, LC, LK, LR, LS, LT, LU, LY, MA, MD, ME, MG, MK, MN, MW, MX, MY, MZ, NA, NG, NI, NO, NZ, OM, PE, PG, PH, PL, PT, QA, RO, RS, RU, RW, SC, SD, SE, SG, SK, SL, SM, ST, SV, SY, TH, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, ZA, ZM, ZW.
- (84) **Designated States (unless otherwise indicated, for every kind of regional protection available):** ARIPO (BW, GH, GM, KE, LR, LS, MW, MZ, NA, RW, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European (AL, AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HR, HU, IE, IS, IT, LT, LU, LV, MC, MK, MT, NL, NO, PL, PT, RO, RS, SE, SI, SK,

[Continued on next page]

(54) **Title:** ATROPISOMERIC 1,8-BISPHENOLNAPHTHALENES AND THEIR USE IN ENANTIOSELECTIVE RECOGNITION AND ASYMMETRIC SYNTHESIS

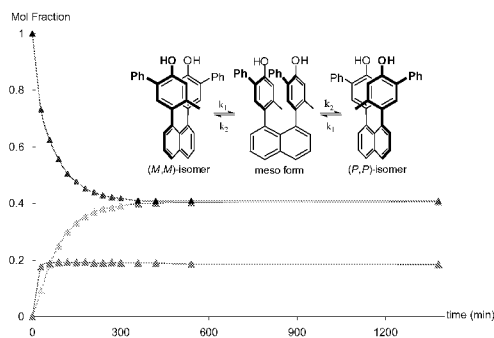
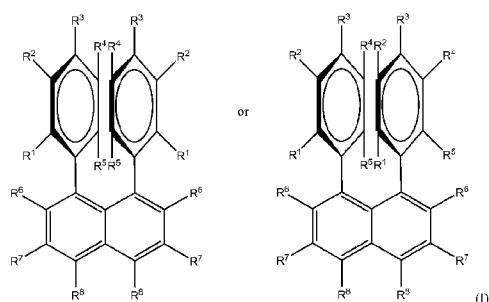


Figure 12



(57) **Abstract:** This invention related to atropisomeric 1,8-bisphenol-naphthalenes and derivatives thereof of the general formula (I): which are useful in resolution of enantiomers, enantioselective recognition and asymmetric synthesis.



SM, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG). **Published:**

— *without international search report and to be republished
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5 **ATROPISOMERIC 1,8-BISPHENOLNAPHTHALENES AND THEIR USE IN**
ENANTIOSELECTIVE RECOGNITION AND ASYMMETRIC SYNTHESIS

STATEMENT OF PRIORITY

This application claims priority to U.S. Provisional Appl. No. 61/471,450, filed
 10 April 4, 2011, the disclosure of which is hereby incorporated by reference.

STATEMENT OF GOVERNMENT INTEREST

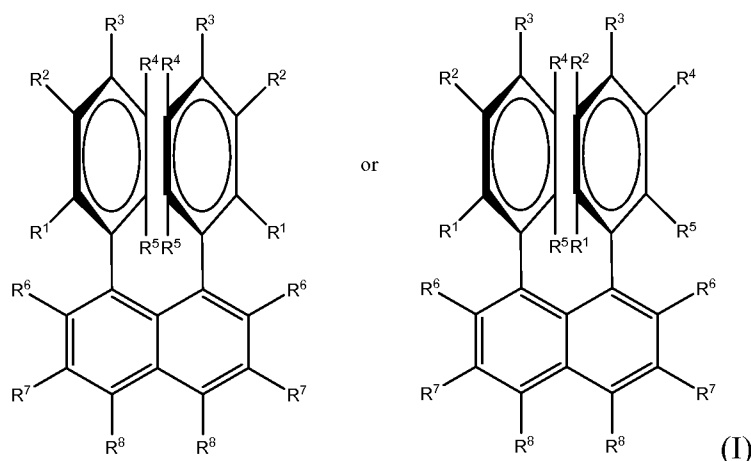
This invention was made in part with government support under grant number
 CHE-0910604, awarded by the National Science Foundation. The government has
 15 certain rights to this invention.

INCORPORATION BY REFERENCE

The documents cited or referenced herein ("herein cited documents"), and all
 documents cited or referenced in herein cited documents, together with any
 20 manufacturer's instructions, descriptions, product specifications, and product sheets for
 any products mentioned herein or in any document incorporated by reference herein, are
 hereby incorporated by reference, and may be employed in the practice of the invention.

FIELD OF THE INVENTION

25 This invention related to atropisomeric 1,8-bisphenolnaphthalenes and
 derivatives thereof of the general formula (I):

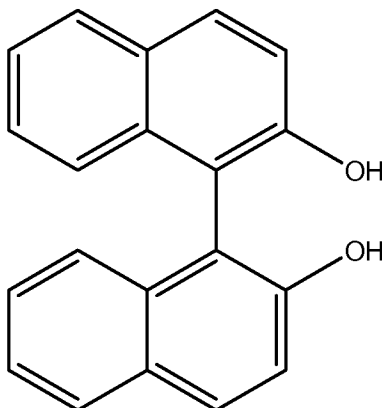


which are useful in resolution of enantiomers, enantioselective recognition and
 asymmetric synthesis.

5 BACKGROUND OF THE INVENTION

The intriguing structure and dynamic stereochemistry of axially chiral compounds has fueled their use in asymmetric synthesis, chiral recognition, the design of microscopic devices such as molecular motors and switches, drug discovery and other areas. (Wolf, C. (Ed.) *Dynamic Stereochemistry of Chiral Compounds*, RSC, Cambridge, 2008.) Undoubtedly, the exceptional diversity and the unique stereochemical, electronic, and photochemical properties of both conformationally stable and rapidly racemizing axially chiral biaryls and polyaryls have led to a wide variety of applications. (See e.g., Qiao, X.; Padula, M. A.; Ho, D. M.; Vogelaar, N. J.; Schutt, C. E.; Pascal Jr., R. A. *J. Am. Chem. Soc.* **1996**, *118*, 741-745.) It is therefore not surprising that structural analysis along with the study of enantiomerization and diastereomerization processes of mono- and disubstituted naphthalenes have received significant attention. (See e.g., Casarini, D.; Lunazzi, L.; Macciantelli, D. *Tetrahedron Lett.* **1984**, *25*, 3641-3642). Alkyl, (See e.g., Fields, D. L.; Regan, T. H. *J. Org. Chem.* **1971**, *36*, 2986-2990.) aryl (See, e.g. House, H. O.; Magin, R. W.; Thompson, H. W. *J. Org. Chem.* **1963**, *28*, 2403-2406.) and heteroaryl (See, e.g. Zoltewicz, J. A.; Maier, N. M.; Fabian, W. M. F. *Tetrahedron* **1996**, *52*, 8703-8706.) groups have been introduced into the naphthalene framework to study the energy barrier to rotation about the naphthyl-alkyl or naphthyl-aryl bond and intramolecular interactions between proximate alkyl and aryl groups.

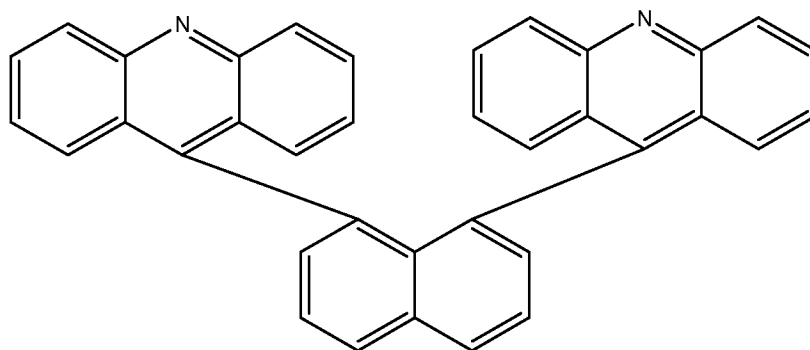
In particular, the incorporation of two phenol rings into a rigid C₂-symmetric scaffold that is reminiscent of the successful BINOL motif has been of general interest due to potential applications in asymmetric catalysis and enantioselective sensing for a long time. (See e.g., Pritchard, R. G.; Steele, M.; Watkinson, M.; Whiting, A. *Tetrahedron Lett.* **2000**, *41*, 6915-6918.)



1,1'-binaphthyl-2,2'-diol (BINOL)

5 Although the synthesis of BINOL was first reported in 1873 it took another 100 years until this prime example of a C_2 -symmetric bidentate atropisomer gained considerable attention. (See von Richter, V. *Chem. Ber.* **1873**, 6, 1249-1260.) Since the mid 1970s, BINOL and its derivatives have found extensive use in asymmetric reactions, molecular recognition studies and other applications. (See e.g., Pu, L. *Chem. Rev.* **1998**, 98, 2405-2494.) The intriguing structure of BINOL and its success as chiral ligand and reagent in asymmetric synthesis has propelled the development of countless analogues that vary in stereoelectronic properties and bite angle. For many years, the synthesis of axially chiral 1,8-bisphenolnaphthalenes has been pursued due to the apparent structural analogy to BINOL and the associated promise in asymmetric catalysis and other areas.

15 Other attempts to synthesize derivatives of BINOL include the applicants previous work described in U.S. Patent 7,888,509 which included general references to aryl and heteroaryl derivatives of naphthalene and specific examples directed to diacridine derivatives of naphthalene (core structure of diacridine derivative shown below).



20 However, the incorporation of sufficient steric bulk into the chiral 1,8-bisphenolnaphthalene framework to halt rotation about the aryl-aryl axes and concomitant racemization has proven difficult. (See e.g., Pritchard, R. G.; Steele, M.; Watkinson, M.; Whiting, A. *Tetrahedron Lett.* **2000**.) Accordingly, few stereodynamic 1,8-bisphenolnaphthalenes such as 1,8-bis(3'-formyl-4'-hydroxyphenyl)naphthalene, **1** have been reported to date and used in racemic form. (See e.g., Watkinson, M.; Whiting, A.; McAuliffe, C. A. *J. Chem. Soc., Chem. Commun.* **1994**, 2141.)

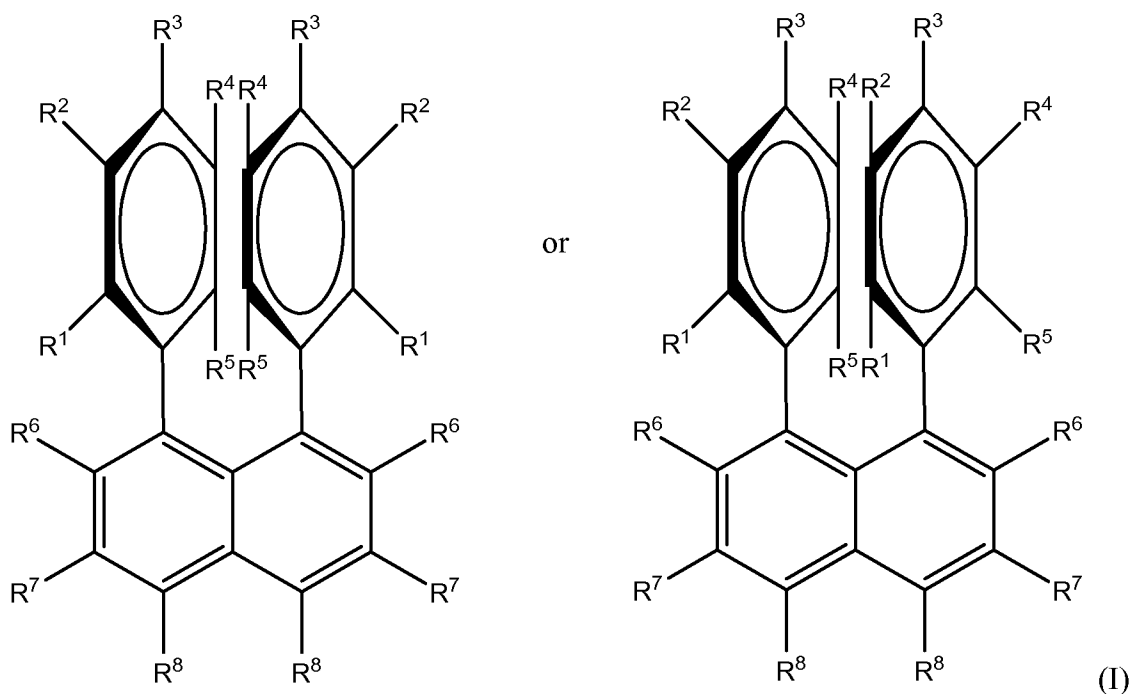
 Therefore, a need still exists in the art for 1,8-bisphenolnaphthalene derivatives which are isolatable and suitable for use in resolution of enantiomers, enantioselective recognition and asymmetric synthesis.

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SUMMARY OF THE INVENTION

5 The applicants have addressed the need in the art for 1,8-bisphenolnaphthalene derivatives which are isolatable and suitable for use in enantioselective recognition and asymmetric synthesis by providing 1,8-bisphenolnaphthalene compounds with ortho and meta substitution on the phenolic ring relative to its bonding with naphthalene.

One aspect of the invention relates to atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I):



wherein R^1 , R^2 , R^3 , R^4 , R^5 , R^6 , R^7 and R^8 are described herein.

Another aspect of the invention relates to the process of making the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I).

15 Another aspect of the invention relates to the process of converting one conformer of the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) into another conformeric form and optionally, to isolate the conformer.

Another aspect of the invention relates to a method of providing enantiomeric recognition between stereoisomers of a compound which comprising of adding an atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I).

Another aspect of the invention relates to the process of separating enantiomers from a racemic mixture of atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I).

25 Another aspect of the invention relates to a method of conducting asymmetric synthesis of a compound which comprising of adding an atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I).

5 These aspects of the invention and other aspects of the invention are disclosed or are apparent from and encompassed by, the Detailed Description which follows below.

 For the purposes of this application the following terms have the meanings defined below:

 “atropisomer” – stereoisomers resulting from hindered rotation about single bonds where
10 the steric strain barrier to rotation is high enough to allow for the isolation of conformers.

 “de” – diastereomeric excess

 “hydroxy protecting group” – refers to known -OH protecting groups to those of skill in the art which includes, but is not limited to those described in *Protective Groups in Organic Synthesis (Fourth Edition)*, Theodora W. Greene and Peter G.M. Wuts, Wiley-
15 Interscience (October 2006).

 For the compound of formula (I) and all subsequent formulae, terms for chemical radicals are used are defined as follows.

 A **hydrocarbon radical** is an aliphatic, cycloaliphatic or aromatic monocyclic or, in the case of an optionally substituted hydrocarbon radical, also a bicyclic or polycyclic
20 organic radical based on the elements carbon and hydrogen, including, for example, the radicals alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, aryl, phenyl, naphthyl, indanyl, indenyl, etc.; this applies correspondingly to hydrocarbon radicals in composite meanings, such as hydrocarbonoxy radicals or other hydrocarbon radicals attached via heteroatom groups.

25 The hydrocarbon radicals, also in the special radicals alkyl, alkoxy, haloalkyl, haloalkoxy, alkylamino and alkylthio, and also the corresponding unsaturated and/or substituted radicals may in each case be straight-chain or branched in the carbon skeleton.

 By way of example, the expression “(C₁-C₄)-alkyl” is a brief notation for alkyl
30 having from 1 to 4 carbon atoms, i.e. encompasses the methyl, ethyl, 1-propyl, 2-propyl, 1-butyl, 2-butyl, 2-methylpropyl or tert-butyl radicals. General alkyl radicals with a larger specified range of carbon atoms, for example “(C₁-C₆)-alkyl” correspondingly also include straight-chain or branched alkyl radicals having a larger number of carbon atoms, i.e., according to the example, also the alkyl radicals having 5 and 6 carbon atoms.

35 **Alkyl** radicals, including in the combined definitions such as alkoxy, haloalkyl, etc., are, for example, methyl, ethyl, n- or i-propyl, n-, i-, t- or 2-butyl, pentyls, hexyls such as n-hexyl, i-hexyl and 1,3-dimethylbutyl, heptyls such as n-heptyl, 1-methylhexyl and 1,4-dimethylpentyl; alkenyl and alkynyl radicals are defined as the possible

5 unsaturated radicals corresponding to the alkyl radicals; alkenyl is, for example, vinyl, allyl, 1-methyl-2-propenyl, 2-methyl-2-propenyl, 2-butenyl, pentenyl, 2-methylpentenyl or hexenyl group, preferably allyl, 1-methylprop-2-en-1-yl, 2-methylprop-2-en-1-yl, but-2-en-1-yl, but-3-en-1-yl, 1-methylbut-3-en-1-yl or 1-methylbut-2-en-1-yl.

Alkenyl also includes in particular straight-chain or branched hydrocarbon
10 radicals having more than one double bond, such as 1,3-butadienyl and 1,4-pentadienyl, but also allenyl or cumulenyl radicals having one or more cumulated double bonds, for example allenyl (1,2-propadienyl), 1,2-butadienyl and 1,2,3-pentatrienyl.

Alkynyl is, for example, propargyl, but-2-yn-1-yl, but-3-yn-1-yl, 1-methylbut-3-yn-1-yl.

15 Alkynyl also includes, in particular, straight-chain or branched hydrocarbon radicals having more than one triple bond or else having one or more triple bonds and one or more double bonds, for example 1,3-butatrienyl or 3-penten-1-yn-1-yl (pent-3-en-1-yn-1-yl).

A **carbocyclic ring** is any form of a closed ring of carbon atoms and can include
20 alicyclic or aromatic structures. Examples of such structures include, but are not limited to (C₃-C₉)-cycloalkyl which is a carbocyclic saturated ring system having 3-9 ring carbon atoms, for example cyclopropyl, cyclobutyl, cyclopentyl, cyclohexyl, cycloheptyl, cyclooctyl or cyclononyl. In the case of substituted cycloalkyl, cyclic systems with substituents are included, where the substituents may also be bonded by a
25 double bond on the cycloalkyl radical, for example an alkylidene group such as methyldiene. Further examples include, but are not limited to (C₅-C₉)-cycloalkenyl which is a carbocyclic, nonaromatic, partially unsaturated ring system having 5-9 ring carbon atoms, for example 1-cyclobutenyl, 2-cyclobutenyl, 1-cyclopentenyl, 2-cyclopentenyl, 3-cyclopentenyl, or 1-cyclohexenyl, 2-cyclohexenyl, 3-cyclohexenyl, 1,3-
30 cyclohexadienyl or 1,4-cyclohexadienyl. In the case of substituted cycloalkenyl, the explanations for substituted cycloalkyl apply correspondingly.

Alkylidene, for example also in the form of (C₁-C₁₀)-alkylidene, is the radical of a straight-chain or branched alkane which is bonded via a double bond, the position of the binding site not being fixed. In the case of a branched alkane, of course, only
35 positions at which two hydrogen atoms may be replaced by the double bond are possible; radicals are, for example, =CH₂, =CH-CH₃, =C(CH₃)-CH₃, =C(CH₃)-C₂H₅ or =C(C₂H₅)-C₂H₅.

Halogen is, for example, fluorine, chlorine, bromine or iodine. Haloalkyl, -

5 alkenyl and -alkynyl are, respectively, alkyl, alkenyl and alkynyl substituted partly or fully by identical or different halogen atoms, preferably from the group of fluorine, chlorine and bromine, in particular from the group of fluorine and chlorine, for example monohaloalkyl, perhaloalkyl, CF₃, CHF₂, CH₂F, CF₃CF₂, CH₂FCHCl, CCl₃, CHCl₂, CH₂CH₂Cl; haloalkoxy is, for example, OCF₃, OCHF₂, OCH₂F, CF₃CF₂O, OCH₂CF₃
10 and OCH₂CH₂Cl; the same applies to haloalkenyl and other halogen-substituted radicals.

Aryl is a mono-, bi- or polycyclic carbocyclic aromatic system, for example phenyl, naphthyl, tetrahydronaphthyl, indenyl, indanyl, pentalenyl, fluorenyl and the like, preferably phenyl.

Optionally substituted aryl also includes polycyclic systems, such as
15 tetrahydronaphthyl, indenyl, indanyl, fluorenyl, biphenyl, where the point of attachment is at the aromatic system.

A **heterocyclic** radical (heterocyclyl) comprises at least one heterocyclic ring (=carbocyclic ring in which at least one carbon atom is replaced by a heteroatom, e.g. by a heteroatom from the group consisting of N, O, S, P, B, Si, Se), which is saturated,
20 unsaturated or heteroaromatic and may be unsubstituted or substituted, where the point of attachment is located at a ring atom.

If the heterocyclyl radical or the heterocyclic ring is optionally substituted, it can be fused to other carbocyclic or heterocyclic rings.

Optionally substituted heterocyclyl also includes polycyclic systems, such as, for
25 example, 8-aza-bicyclo[3.2.1]octanyl or 1-aza-bicyclo[2.2.1]heptyl.

Optionally substituted heterocyclyl also includes spirocyclic systems, such as, for example, 1-oxa-5-aza-spiro[2.3]hexyl.

Heteroaryl means, from among the systems defined above under "heterocyclyl", in each case a heteroaromatic compound, i.e. a fully unsaturated aromatic heterocyclic
30 compound.

Substituted radicals, such as a substituted alkyl, alkenyl, alkynyl, cycloalkyl, aryl, phenyl, benzyl, heterocyclyl and heteroaryl radical, are, for example, a substituted radical derived from the unsubstituted base structure, where the substituents are, for example, one or more, preferably 1, 2 or 3, radicals selected from the group of halogen,
35 alkoxy, alkylthio, SF₅, hydroxyl, amino, nitro, carboxyl, cyano, azido, alkoxycarbonyl, alkylcarbonyl, formyl, carbamoyl, mono- and dialkylaminocarbonyl, substituted amino such as acylamino, mono- and dialkylamino, and alkylsulfinyl, alkylsulfonyl and, in the case of cyclic radicals, also alkyl, haloalkyl, alkylthioalkyl, alkoxyalkyl, optionally

5 substituted mono- and dialkylaminoalkyl and hydroxyalkyl; in the term "substituted radicals", such as the substituted alkyl, alkenyl, alkynyl, cycloalkyl, aryl, phenyl, benzyl, heterocyclyl and heteroaryl radicals mentioned, substituents include, in addition to the saturated hydrocarbon radicals mentioned such as alkyl, alkoxy, alkylthio, alkoxy carbonyl, haloalkyl, cycloalkyl or cycloalkyloxy, corresponding unsaturated
10 aliphatic and aromatic radicals, such as optionally substituted alkenyl, alkynyl, alkenyloxy, alkynyloxy, phenyl and phenoxy. In the case of substituted cyclic radicals having aliphatic moieties in the ring, cyclic systems with those substituents which are bonded on the ring by a double bond are also included, for example substituted by an alkylidene group such as methylenidene or ethylenidene.

15 **Acyl** is a radical of an organic acid which arises in a formal sense by removal of a hydroxyl group on the acid function, and the organic radical in the acid may also be bonded to the acid function via a heteroatom. Examples of acyl are the -CO-R radical of a carboxylic acid HO-CO-R and radicals of acids derived therefrom, such as those of thiocarboxylic acid, optionally N-substituted iminocarboxylic acids or the radical of
20 carbonic monoesters, N-substituted carbamic acid, sulfonic acids, sulfinic acids, N-substituted sulfonamide acids, phosphonic acids or phosphinic acids.

The "yl position" of a radical denotes the carbon atom having the free bond. Compounds of the formula (I) according to the invention and compounds of the formula (I) used according to the invention (and, if appropriate, salts thereof) are in short also
25 referred to as "compounds (I)".

In addition, for the purposes of this application compounds of the invention also include all stereoisomers and racemic mixtures thereof. The possible stereoisomers defined by their specific three-dimensional shape, such as enantiomers, diastereomers, Z- and E-isomers, are all encompassed by the formula (I) and can, in the preferred
30 enantioselective procedure, be prepared selectively when optically active materials are used. The application also includes all tautomeric forms, pharmaceutically acceptable salts and crystalline forms including polymorphic forms.

BRIEF DESCRIPTION OF THE DRAWINGS

35 The following detailed description, given by way of example, but not intended to limit the invention solely to the specific embodiments described, may best be understood in conjunction with the accompanying drawings, in which:

5 Figure 1 depicts ^1H NMR spectra of diimine diastereomers obtained from racemic **5** and (*R*)-2-amino-1-propanol and quantitative conversion towards (*P,P,R,R*)-**9**. (A) 60 °C, 0 h, (B) 60.0 °C, 14 h. * corresponds to (*M,M,R,R*)-**9**, ‡ corresponds to (*P,P,R,R*)-**9**.

Figure 2 depicts a CD spectra of (*P,P*)-**5** (blue), (*P,P,R,R*)-**9** (red) and (*P,P,S,S*)-**9** (green) at 5.0×10^{-5} M in CHCl_3 .
10

Figure 3 depicts different views of the crystal structure of (*P,P*)-**5**.

Figure 4 depicts a decrease of the CD signal of (*P,P*)-**5** as a result of racemization in chloroform (6.95×10^{-4} M) at 45.0 °C. The CD spectra were collected at 25 °C at a concentration of 5.0×10^{-5} M in chloroform.

15 Figure 5 depicts a change in the mole fraction of (*P,P*)-**5** upon heating to 45.0 °C. For conditions, see Figure 4.

Figure 6 depicts the analysis and curve fitting of the change in the mol fractions of (*P,P,S,S*)-**9** (red), (*M,P,S,S*)-**9** (black) and (*M,M,S,S*)-**9** (blue) upon heating of (*P,P,S,S*)-**9** to 58.0 °C in chloroform.

20 Figure 7 depicts the single crystal structure of *syn*-**9**.

Figure 8 depicts different views of the crystal structure of (*P,P,R,R*)-**10** showing the hydrogen bonding motif.

Figure 9 depicts a CD spectra of (*P,P*)-**5** (blue), (*P,P,R,R*)-**10** (red) at 5.0×10^{-5} M in CHCl_3 .

25 Figure 10 depicts the change in the CD signal of (*P,P,S,S*)-**10** as a result of diastereomerization at 50.0 °C. The CD spectra were collected at 25.0 °C with a concentration of 5.0×10^{-5} M in CHCl_3 .

Figure 11 depicts a plot of the mol fraction of (*P,P,S,S*)-**9** versus time (min).

Figure 12 depicts the change in the mol fractions of (-)-**2** (blue), (+)-**2** (orange)
30 and *syn*-**2** (green) during heating at 77.1 °C.

Figure 13 depicts of a HPLC chromatogram of racemic **2**.

Figure 14a depicts a UV spectrum of **2** (1.02×10^{-3} M, hexanes:IPA 1:1).

Figure 14b depicts a CD spectra of dextrorotatory (blue) and levorotatory (orange) enantiomers of **2** (9.45×10^{-5} M, ACN).

35 Figure 15 depicts CD Spectra of dextrorotatory (blue) and levorotatory (orange) enantiomers of **2** (9.45×10^{-5} M, ACN). The dashed lines show the CD spectra upon treatment with sodium *tert*-butoxide.

5 Figure 16 depicts the change of the %ee of **2** during heating at 77.1 °C. Inset: HPLC chromatograms showing the change in the relative amounts of the stereoisomers of **2**.

 Figure 17 depicts the change in the mol fractions of (-)-**2** (blue), (+)-**2** (orange) and *syn*-**2** (green) during heating at 77.1 °C.

10 Figure 18 depicts the single crystal structure of *anti*-**2**.

 Figure 19 depicts different views of the crystal structure of **2**.

 Figure 20 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R*)-1-phenylethylamine (blue) and the (*S*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

15 Figure 21 depicts a change in the CD spectrum of the bisphenoxides of (+)-**2** (right) upon addition of (*R*)-1-phenylethylamine (blue) and the (*S*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

 Figure 22 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R*)-3,3-dimethylbutan-2-amine (blue) and the (*S*)-enantiomer (red). The
20 concentration of **2** was 9.45×10^{-5} M in acetonitrile.

 Figure 23 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R*)-3-dimethylbutan-2-amine (blue) and the (*S*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

 Figure 24 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2**
25 upon addition of (*R,R*)-cyclohexane-1,2-diamine (blue) and the (*S,S*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

 Figure 25 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R,R*)-1,2-diphenylethane-1,2-diamine (blue) and the (*S,S*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

30 Figure 26 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R*)-1,1-diphenylpropan-2-amine (blue) and the (*S*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

 Figure 27 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R*)-1-cyclohexylethanamine (blue) and the (*S*)-enantiomer (red). The
35 concentration of **2** was 9.45×10^{-5} M in acetonitrile.

 Figure 28 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R,R,R,S*)-2,6,6-trimethylbicyclo[3.1.1]heptan-3-amine (blue) and the (*S,S,S,R*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

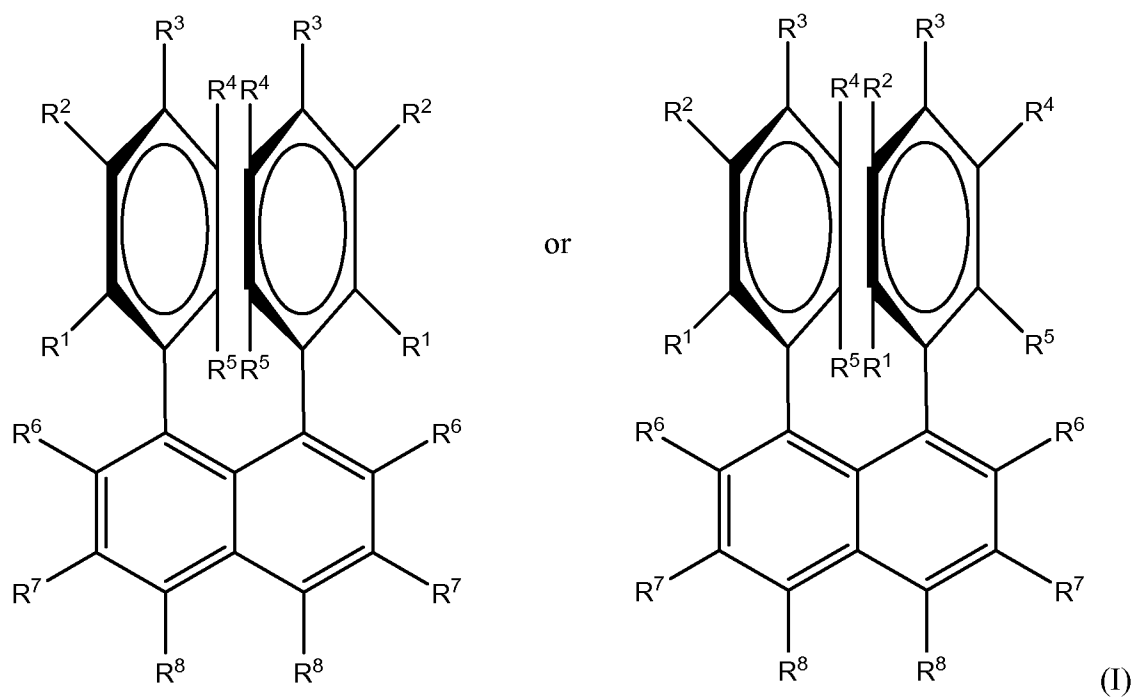
5 Figure 29 depicts a change in the CD spectrum of the bisphenoxides of (-)-**2** upon addition of (*R*)-heptan-2-amine (blue) and the (*S*)-enantiomer (red). The concentration of **2** was 9.45×10^{-5} M in acetonitrile.

Figure 30 depicts a change in the CD spectrum of the bisphenoxides of (+)-**2** (right) upon addition of (*R*)-heptan-2-amine (blue) and the (*S*)-enantiomer (red). The
10 concentration of **2** was 9.45×10^{-5} M in acetonitrile.

DETAILED DESCRIPTION

One aspect of the invention relates to atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I).

15 One embodiment of this aspect of the invention is the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I):



wherein

20 R^1 and R^5 are independently selected from the group consisting of hydrogen, alkyl, alkenyl, alkynyl, cycloalkyl, aryl, benzyl, heterocyclyl, heteroaryl, chloro, bromo, iodo, acyl, amino, amido, azido, cyano, formyl, carbamoyl, $-SF_5$, nitro, $-OR'$, $-NR''R'''$, $-SR''''$, $-P(O)(OR^A)$, $-P(OR^A)_2$, and $P(R^B)_2$,

wherein

25 R' is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical,

5 heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical, -
 $P(O)(OR^A)$, $-P(OR^A)_2$, and $P(R^B)_2$; or OR' is an optionally substituted saturated or
 unsaturated, nonaromatic or aromatic heterocyclic radical;

10 R'' and R''' are each independently of one another H, alkyl, alkenyl, alkynyl,
 cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical,
 hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical,
 heterocyclylthio radical or heterocyclylamino radical; or $NR''R'''$ is an optionally
 substituted saturated or unsaturated, nonaromatic or aromatic heterocyclic
 radical;

15 R'''' are each independently of one another H, alkyl, alkenyl, alkynyl, cycloalkyl,
 cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical,
 a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or
 heterocyclylamino radical; or SR'' is an optionally saturated or unsaturated,
 20 nonaromatic or aromatic heterocyclic radical;

R^A is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic)
 hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical,
 heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

25 R^B is H, halogen, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or
 cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical,
 heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

wherein at least one of R^1 and R^5 on each phenyl ring is not hydrogen;

30 R^2 , R^3 , R^4 , R^6 , R^7 and R^8 are independently selected from the group consisting of
 hydrogen, alkyl, alkenyl, alkynyl, cycloalkyl, aryl, benzyl, heteroalkyl, heterocyclyl,
 heterocycloalkyl, heteroaryl, halogen, acyl, amino, amido, azido, alkyliminyl ($-C=NH-$
 alkyl), carboxy, cyano, formyl, carbamoyl, $-SF_5$, nitro, OR' , $NR''R'''$, SR'''' , -

35 $P(O)(OR^A)$, $-P(OR^A)_2$, and $P(R^B)_2$,

wherein

R' is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic)
 hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical,

5 heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical, -
 $P(O)(OR^A)$, $-P(OR^A)_2$, and $P(R^B)_2$; or OR' is an optionally substituted saturated or
 unsaturated, nonaromatic or aromatic heterocyclic radical;

10 R'' and R''' are each independently of one another H, alkyl, alkenyl, alkynyl,
 cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical,
 hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical,
 heterocyclylthio radical or heterocyclylamino radical; or $NR''R'''$ is an optionally
 substituted saturated or unsaturated, nonaromatic or aromatic heterocyclic
 radical;

15 R'''' are each independently of one another H, alkyl, alkenyl, alkynyl, cycloalkyl,
 cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical,
 a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or
 heterocyclylamino radical; or SR'' is an optionally saturated or unsaturated,
 20 nonaromatic or aromatic heterocyclic radical;

R^A is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic)
 hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical,
 heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

25 R^B is H, halogen, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or
 cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical,
 heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

wherein at least one of R^2 and R^4 on each phenyl ring is not hydrogen.

30

Another embodiment of this aspect of the invention is the atropisomeric 1,8-
 bisphenolnaphthalene compounds of formula (I), wherein

R^1 and R^5 are independently selected from the group consisting of hydrogen,
 (C₁-C₆) alkyl, (C₂-C₆) alkenyl, (C₂-C₆) alkynyl, (C₃-C₉) cycloalkyl, aryl, benzyl,
 35 heterocyclyl, heteroaryl, chloro, bromo, iodo, (C₁-C₆) acyl, amino, amido, azido, cyano,
 formyl, carbamoyl, -SF₅, nitro, OR' , $NR''R'''$, SR'''' , $-P(O)(OR^A)$, $-P(OR^A)_2$, and
 $P(R^B)_2$,
 wherein

5 R' is H, (C₁-C₆) alkyl, (C₂-C₆) alkenyl, (C₂-C₆) alkynyl, (C₃-C₆) cycloalkyl or (C₅-C₆) cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical or hydrocarbonoxy radical having in each case 1 to 10 carbon atoms, preferably 1 to 6 carbon atoms, or a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical having in each case 3 to 9
10 ring atoms and 1 to 3 hetero ring atoms from the group consisting of N, O and S, where each of the six last-mentioned radicals is unsubstituted or substituted, or an acyl radical of an organic acid having 1 to 20 carbon atoms, preferably 1 to 12 carbon atoms, in particular 1 to 8 carbon atoms, -P(O)(OR^A), -P(OR^A)₂, and P(R^B)₂; or

15 OR' is a saturated or unsaturated, nonaromatic or aromatic heterocyclic radical having 3 to 9 ring atoms and 1 to 4 hetero ring atoms with at least the oxygen from OR' being a hetero ring atom, where the further hetero ring atoms optionally present in addition to the nitrogen atom are selected from the group consisting of N, O and S and the radical is unsubstituted or substituted; or

20 R'' and R''' are each independently of one another H, (C₁-C₆)alkyl, (C₂-C₆)alkenyl, (C₂-C₆)alkynyl, (C₃-C₆)cycloalkyl or (C₅-C₆)cycloalkenyl, or an (acyclic or cyclic) hydrocarbon radical or hydrocarbonoxy radical having in each case 1 to 10 carbon atoms, preferably 1 to 6 carbon atoms, or a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino
25 radical having in each case 3 to 9 ring atoms and 1 to 3 hetero ring atoms from the group consisting of N, O and S, where each of the six last-mentioned radicals is unsubstituted or substituted, or an acyl radical of an organic acid having 1 to 20 carbon atoms, preferably 1 to 12 carbon atoms, in particular 1 to 8 carbon atoms;
30 or

NR''R''' is a saturated or unsaturated, nonaromatic or aromatic heterocyclic radical having 3 to 9 ring atoms and 1 to 4 hetero ring atoms with at least the nitrogen from NR''R''' being a hetero ring atom, where the further hetero ring atoms optionally present in addition to the nitrogen atom are selected
35 from the group consisting of N, O and S and the radical is unsubstituted or substituted; or

5 R'''' is H, (C₁-C₆)alkyl, (C₂-C₆)alkenyl, (C₂-C₆)alkynyl, (C₃-C₆)cycloalkyl or (C₅-C₆)cycloalkenyl, or an (acyclic or cyclic) hydrocarbon radical or hydrocarbonoxy radical having in each case 1 to 10 carbon atoms, preferably 1 to 6 carbon atoms, or a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical having in each case 3 to 9
 10 ring atoms and 1 to 3 hetero ring atoms from the group consisting of N, O and S, where each of the six last-mentioned radicals is unsubstituted or substituted, or an acyl radical of an organic acid having 1 to 20 carbon atoms, preferably 1 to 12 carbon atoms, in particular 1 to 8 carbon atoms, or

SR'''' is a saturated or unsaturated, nonaromatic or aromatic heterocyclic
 15 radical having 3 to 9 ring atoms and 1 to 4 hetero ring atoms with at least the sulfur from SR' being a hetero ring atom, where the further hetero ring atoms optionally present in addition to the nitrogen atom are selected from the group consisting of N, O and S and the radical is unsubstituted or substituted;

20 R^A is H, or (C₁-C₆) alkyl;

R^B is H, halogen, or (C₁-C₆) alkyl,;

wherein at least one of R¹ and R⁵ on each phenyl ring is not hydrogen;

25 R², R³, R⁴, R⁶, R⁷ and R⁸ are independently selected from the group consisting of hydrogen, (C₁-C₆) alkyl, (C₂-C₆) alkenyl, (C₂-C₆) alkynyl, (C₃-C₉) cycloalkyl, aryl, benzyl, heterocyclyl, heteroaryl, chloro, bromo, iodo, (C₁-C₆) acyl, amino, amido, azido, cyano, formyl, carbamoyl, -SF₅, nitro, OR', NR''R''', SR'''', -P(O)(OR^A), -P(OR^A)₂,
 30 and P(R^B)₂,

wherein

R' is H, (C₁-C₆) alkyl, (C₂-C₆) alkenyl, (C₂-C₆) alkynyl, (C₃-C₆) cycloalkyl or (C₅-C₆) cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical or hydrocarbonoxy radical having in each case 1 to 10 carbon atoms, preferably 1 to 6 carbon atoms, or a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical having in each case 3 to 9
 35 ring atoms and 1 to 3 hetero ring atoms from the group consisting of N, O and S, where each of the six last-mentioned radicals is unsubstituted or substituted, or an

5 acyl radical of an organic acid having 1 to 20 carbon atoms, preferably 1 to 12 carbon atoms, in particular 1 to 8 carbon atoms, $-P(O)(OR^A)$, $-P(OR^A)_2$, and $P(R^B)_2$; or

OR' is a saturated or unsaturated, nonaromatic or aromatic heterocyclic radical having 3 to 9 ring atoms and 1 to 4 hetero ring atoms with at least the oxygen from OR' being a hetero ring atom, where the further hetero ring atoms optionally present in addition to the nitrogen atom are selected from the group consisting of N, O and S and the radical is unsubstituted or substituted; or

15 R'' and R''' are each independently of one another H, (C₁-C₆)alkyl, (C₂-C₆)alkenyl, (C₂-C₆)alkynyl, (C₃-C₆)cycloalkyl or (C₅-C₆)cycloalkenyl, or an (acyclic or cyclic) hydrocarbon radical or hydrocarbonoxy radical having in each case 1 to 10 carbon atoms, preferably 1 to 6 carbon atoms, or a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical having in each case 3 to 9 ring atoms and 1 to 3 hetero ring atoms from the group consisting of N, O and S, where each of the six last-mentioned radicals is unsubstituted or substituted, or an acyl radical of an organic acid having 1 to 20 carbon atoms, preferably 1 to 12 carbon atoms, in particular 1 to 8 carbon atoms; or

25 NR''R''' is a saturated or unsaturated, nonaromatic or aromatic heterocyclic radical having 3 to 9 ring atoms and 1 to 4 hetero ring atoms with at least the nitrogen from NR''R''' being a hetero ring atom, where the further hetero ring atoms optionally present in addition to the nitrogen atom are selected from the group consisting of N, O and S and the radical is unsubstituted or substituted; or

30 R'''' is H, (C₁-C₆)alkyl, (C₂-C₆)alkenyl, (C₂-C₆)alkynyl, (C₃-C₆)cycloalkyl or (C₅-C₆)cycloalkenyl, or an (acyclic or cyclic) hydrocarbon radical or hydrocarbonoxy radical having in each case 1 to 10 carbon atoms, preferably 1 to 6 carbon atoms, or a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical having in each case 3 to 9 ring atoms and 1 to 3 hetero ring atoms from the group consisting of N, O and S, where each of the six last-mentioned radicals is unsubstituted or substituted, or an

5 acyl radical of an organic acid having 1 to 20 carbon atoms, preferably 1 to 12 carbon atoms, in particular 1 to 8 carbon atoms, or

SR'''' is a saturated or unsaturated, nonaromatic or aromatic heterocyclic radical having 3 to 9 ring atoms and 1 to 4 hetero ring atoms with at least the sulfur from SR' being a hetero ring atom, where the further hetero ring atoms
 10 optionally present in addition to the nitrogen atom are selected from the group consisting of N, O and S and the radical is unsubstituted or substituted;

R^A is H, or (C₁-C₆) alkyl;

R^B is H, halogen, or (C₁-C₆) alkyl;

15 wherein at least one of R² and R⁴ on each phenyl ring is not hydrogen.

Another embodiment of this aspect of the invention is the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I), wherein

R¹ and R⁵ are independently selected from the group consisting of hydrogen, (C₁-C₄)
 20 alkyl, phenyl, chloro, bromo and iodo, wherein at least one of R¹ and R⁵ on each phenyl ring is not hydrogen;

R² and R⁴ are independently selected from the group consisting of hydrogen, (C₁-C₄) alkyl, (C₂-C₄) alkenyl, (C₂-C₄) alkynyl, (C₃-C₆) cycloalkyl, phenyl, (C₁-C₆) acyl, and formyl, wherein at least one of R² and R⁴ on each phenyl ring is not hydrogen.;

25 R³ is independently PR'₂, NR'₂ or OR' wherein each R' is independently H or (C₁-C₄) alkyl, and

R⁶, R⁷ and R⁸ are independently selected from the group consisting of hydrogen, (C₁-C₄) alkyl, hydroxyl, (C₁-C₄) alkoxy, halogen, NH₂, amido, azido, cyano, formyl, carbamoyl, nitro.

30

Another embodiment of this aspect of the invention is the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I), wherein

R¹ and R⁵ are independently selected from the group consisting of hydrogen, (C₁-C₄) alkyl, phenyl, chloro, bromo and iodo, wherein at least one of R¹ and R⁵ on each
 35 phenyl ring is not hydrogen;

R² and R⁴ are independently selected from the group consisting of hydrogen, (C₁-C₄) alkyl, (C₂-C₄) alkenyl, (C₂-C₄) alkynyl, (C₃-C₆) cycloalkyl, phenyl, (C₁-C₆) acyl, and formyl, wherein at least one of R² and R⁴ on each phenyl ring is not hydrogen.;

5 R^3 is independently OR' wherein R' is H or (C_1-C_4) alkyl; and
 R^6 , R^7 and R^8 are independently selected from the group consisting of hydrogen,
 (C_1-C_4) alkyl, hydroxyl, (C_1-C_4) alkoxy, halogen, NH_2 , amido, azido, cyano, formyl,
 carbamoyl, nitro.

10 Another embodiment of this aspect of the invention is the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I), wherein

R^1 and R^5 are independently H and methyl, wherein at least one of R^1 and R^5 on each phenyl ring is not hydrogen;

15 R^2 and R^4 are independently phenyl or formyl, wherein at least one of R^2 and R^4 on each phenyl ring is not hydrogen;

R^3 is independently OR' wherein R' is H or methyl; and

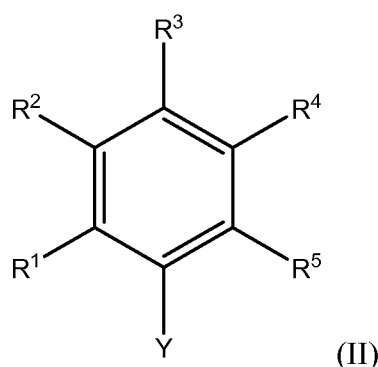
R^6 , R^7 and R^8 are independently selected from the group consisting of hydrogen and methyl.

20 In another aspect of the invention, the compounds of formula (I) are stable to racemization at room temperature (20 - 25°C).

Another aspect of the invention relates to the process of making the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I).

25 One embodiment of this aspect of the invention is where the process comprises of:

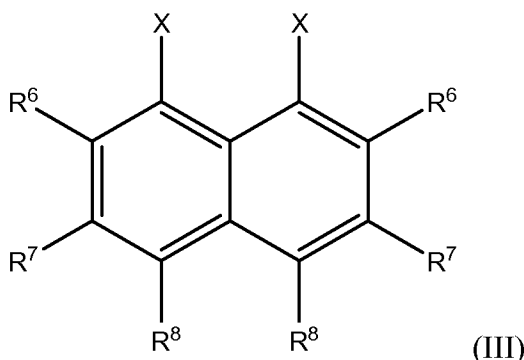
(a) reacting a compound of formula (II):



wherein R^1 , R^2 , R^3 , R^4 and R^5 are as defined above and Y is a boron- or

30 metal-based moiety
 with a compound of formula (III)

19



5

wherein R^6 , R^7 and R^8 are as defined above and X is a displaceable group to form the compounds of formula (I); and

(b) optionally, if R^3 is not hydroxyl after step (a), deprotecting or hydrolyzing R^3 to form the compound of formula (I) wherein R^3 is hydroxyl.

10

In another embodiment of this aspect of the invention, Y is a boron based moiety and X is selected from the group consisting of chlorine, bromine, iodine, phenylsulfonate, tosylate and triflate.

In another embodiment of this aspect of the invention, Y is selected from the group consisting of $B(OH)_2$, 9-BBN, $B(CH_2CH_3)_2$.

15

In another embodiment of this aspect of the invention, is where the process is by way of a process selected from the group consisting of Suzuki coupling, Stille coupling, Negishi coupling and Kumada coupling.

20

In another embodiment of this aspect of the invention, the process results in a ratio of *syn*- and *anti*-isomers for the compounds of formula (I) selected from the ranges consisting of about 1:1 to about 1:8, about 1:2 to about 1:6 and about 1:3 to about 1:4

In another embodiment of this aspect of the invention, the process results in % de (% diastereomeric excess) for the compounds of formula (I) selected from the ranges consisting of >95%, >98% and >99%. Note that >99% is effectively complete stereochemical purity

25

In another embodiment of this aspect of the invention, the moieties represented by variables R^6 - R^8 can be synthesized after forming the the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) with the above definition of variables R^1 - R^5 by derivitization techniques known to those of skill in the art (see e.g. *March's Advanced Organic Chemistry – Reactions, Mechanisms and Structure (6th Edition)*, Michael B. Smith and Jerry March, Wiley-Interscience (2007); *Modern Synthetic Reactions (Second Edition)*, Herbert House, W.A. Benjamin, Inc. (1972)).

30

5 Another aspect of the invention relates to the process of converting one conformer of the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) into another conformeric form.

In one embodiment of this aspect of the invention, a conformer of the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) obtained from the described process of making is converted to another conformer by dissolving the
10 conformer of the compound of formula (I) in a solvent and heating. In another embodiment of the invention, the heating can be performed from room temperature to a temperature range selected from the group consisting of about 35°C to less than the boiling point of the solvent; about 35°C to about 95°C and about 40°C to about 65°C.
15 To facilitated the heating and conversion process, the compounds of formula (I) can be dissolved using known techniques in the art, e.g. such as those described in *Vogel's Textbook of Practical Organic Chemistry (Fifth Edition)*, Furniss et al., Longman Scientific & Technical (1989)).

Another aspect of the invention relates to a process of separating enantiomers
20 from a racemic mixture of atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I).

In one embodiment of this aspect of the invention, a racemic mixture of atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) is allowed to react with a chiral compound to form a diastereomeric adduct.

25 In another embodiment of this aspect of the invention, R^2 and R^4 of formula (I) are independently selected from the group consisting of carboxy, amide, -CHO, - (CO)alkyl, (C=NH)alkyl and amidino.

In another embodiment of this aspect of the invention, the chiral compound is an amine, an amide, a carboxylic acid, an amino alcohol or an alcohol.

30 In another embodiment of this aspect of the invention, the chiral compound is an amine or an amino alcohol.

In another embodiment of this aspect of the invention, the chiral compound is an amino alcohol.

In another embodiment of this aspect of the invention, diastereomeric adducts are
35 formed by condensing the racemic mixture of atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) with a chiral amine or amino alcohol to form a diimine product.

5 In another embodiment of this aspect of the invention, the chiral amino alcohol is selected from the group consisting of 2-amino-1-propanol, 2-amino-4-methyl-1-pentanol, ephedrine or pseudoephedrine.

In another embodiment of this aspect of the invention, the mixture of diastereomeric adducts are heated to allow the interconversion between the
10 diastereomers, which results in a diastereomeric adduct of formula (I) with a % diasteriomic excess (% de) selected from the ranges consisting of >90%, >95%, >98% and >99%.

In another embodiment of this aspect of the invention, the diastereomers of the diastereomeric adduct of formula I are separated chromatographically, which results in a
15 diastereomeric adduct of formula I with a % diasteriomic excess (% de) selected from the ranges consisting of >90%, >95%, >98% and >99%.

In another embodiment of this aspect of the invention, the diasteriomic adduct is cleaved to yield an atropisomeric 1,8-bisphenolnaphthalene compound of formula (I) with % enantiomeric excess (% ee) selected from the ranges consisting of >90%, >95%,
20 >98% and >99%. Note that >99% is effectively complete stereochemical purity.

In another embodiment of this aspect of the invention, the cleavage of the diasteriomic adduct is accomplished by hydrolysis to yield an atropisomeric 1,8-bisphenolnaphthalene compound of formula (I) with % enantiomeric excess (% ee) selected from the ranges consisting of >90%, >95%, >98% and >99%.

25 Another aspect of the invention relates to a method of providing enantiomeric recognition between stereoisomers of a chiral compound which comprises of adding an atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) to a solution containing a racemic or diasteriomic mixture of the chiral compound..

30 In one embodiment of this aspect of the invention, the chiral compound is an amine, amino alcohol amino acid or alcohol.

In another embodiment of the invention, the chiral compound is an amine.

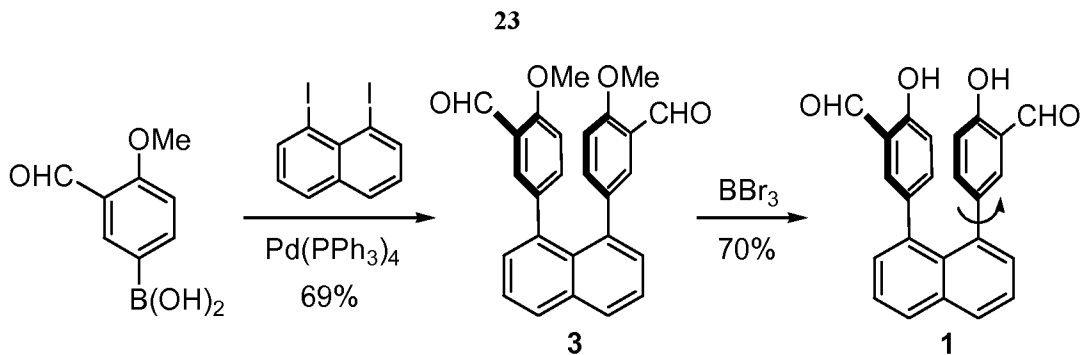
In another embodiment of the invention, the chiral amine is selected from the group consisting of 1-phenylethylamine, 3,3-dimethylbutan-2-amine, 3-dimethylbutan-2-amine, cyclohexane-1,2-diamine, 1,2-diphenylethane-1,2-diamine, 1,1-diphenylpropan-
35 2-amine, 1-cyclohexylethanamine, 2,6,6-trimethylbicyclo[3.1.1]heptan-3-amine, and heptan-2-amine.

5 In another embodiment of this aspect of the invention, the atropisomeric 1,8-bisphenolnaphthalene compounds of formula (I) is **2**.

 Another aspect of the invention relates to a method of conducting asymmetric synthesis of a compound which comprising of adding an atropisomeric 1,8-
10 bisphenolnaphthalene compounds of formula (I).

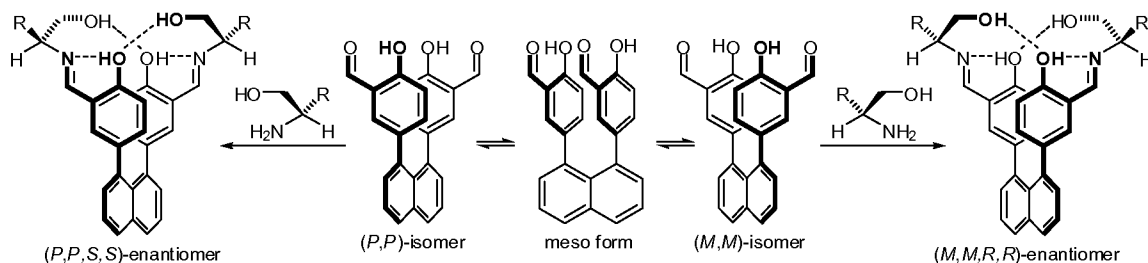
 Further details related to the aspects and embodiments of the invention follow below.

15 In continuation of previously conducted studies with stereodynamic chiral biaryls and triaryls, (See (a) Wolf, C.; Ghebramariam, B. T. *Tetrahedron: Asymm.* **2002**, *13*, 1153-1156. (b) Wolf, C.; Tumambac, G. E. *J. Phys. Chem. A.* **2003**, *107*, 815-817. (c) Tumambac, G.E.; Wolf, C. *J. Org. Chem.* **2004**, *69*, 2048-2055. (d) Tumambac, G. E.; Mei, X.; Wolf, C. *Eur. J. Org. Chem.* **2004**, 3850-3856. (e) Wolf, C. *Chem. Soc. Rev.* **2005**, *34*, 595-608. (f) Tumambac, G. E.; Wolf, C. *J. Org. Chem.* **2005**, *70*, 2930-2938. (g) Wolf, C.; Xu, H. *Tetrahedron Lett.* **2007**, *48*, 6886-6889.) and 1,8-diheteroarylnaphthalene-derived sensors, ((a) Mei, X.; Wolf, C. *Chem. Commun.* **2004**, 2078-2079. (b) Mei, X.; Wolf, C. *J. Am. Chem. Soc.* **2004**, *126*, 14736-14737. (c) Tumambac, G. E.; Wolf, C. *Org. Lett.* **2005**, *7*, 4045-4048. (d) Mei, X.; Martin, R. M.;
25 Wolf, C. *J. Org. Chem.* **2006**, *71*, 2854-2861. (e) Liu, S.; Pestano, J. P. C.; Wolf, C. *J. Org. Chem.* **2008**, *73*, 4267-4270. (f) Mei, X.; Wolf, C. *Tetrahedron Lett.* **2006**, *47*, 7901-7904. (g) Wolf, C.; Liu, S.; Reinhardt, B. C. *Chem. Commun.* **2006**, 4242-4244. (h) Mei, X.; Wolf, C. *J. Am. Chem. Soc.* **2006**, *128*, 13326-13327.) 1,8-bis(3'-formyl-4'-hydroxyphenyl)naphthalene, **1**, exhibiting two salicylaldehyde rings in the *peri*-positions
30 of naphthalene was prepared via Suzuki coupling of 1,8-diiodonaphthalene and boronic acid followed by deprotection of dialdehyde **3** (Scheme 1). (See (a) Watkinson, M.; Whiting, A.; McAuliffe, C. A. *J. Chem. Soc., Chem. Commun.* **1994**, 2141. (b) Ghosn, M. W.; Wolf, C. *J. Am. Chem. Soc.* **2009**, *131*, 16360-16361.)



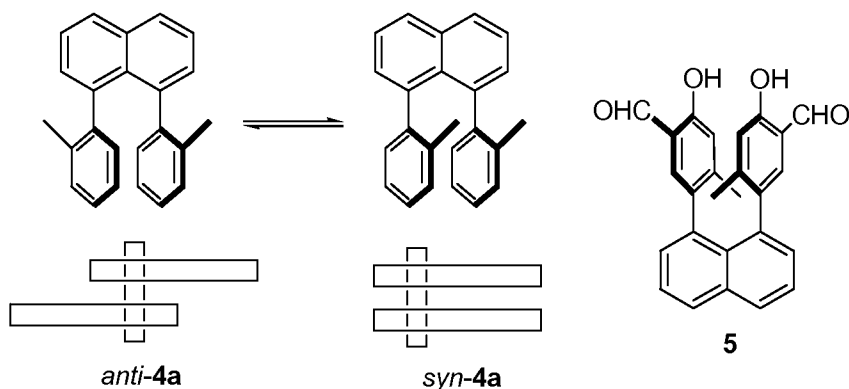
Scheme 1. Synthesis of 1,8-bis(3'-formyl-4'-hydroxyphenyl)naphthalene, **1**.

Triaryl **1** undergoes fast rotation about the two aryl-aryl bonds at room temperature, which results in the interconversion of the enantiomeric *anti*-isomers via the thermodynamically less stable meso *syn*-intermediate. We realized that imine formation with amino alcohols disturbs this equilibrium and strongly favors population of a single diastereomer that is stabilized by intramolecular hydrogen bonding (Scheme 2). The diimine formed displays strong Cotton effects at high wavelengths and NMR and crystallographic analysis showed that the central chirality of the amino alcohol substrate induces a rigid, axially chiral triaryl scaffold with perfect stereocontrol: Condensation of **1** and (*R*)-amino alcohols results in well-defined amplification of asymmetric induction and the triaryl was found to adopt an (*M,M*)-conformation. The opposite sense of chiral induction was observed with (*S*)-amino alcohols. We were able to demonstrate that the fast diimine formation, which is complete within 5 minutes, followed by *in situ* CD measurements allows time-efficient determination of the absolute configuration and the enantiomeric purity of the substrate used. Similar results were obtained with amino acids. (See Ghosn, M. W.; Wolf, C. *Tetrahedron* **2010**, 66, 3989-3994.)



Scheme 2. Central-to-axial chirality induction upon diimine formation with stereodynamic triaryl **1**.

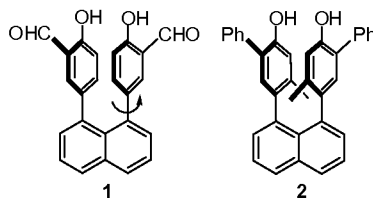
5 We envisioned that a less fluxional analogue of **1** would provide further insights into (a) the amplification of asymmetric induction, (b) the effect of the intramolecular hydrogen bonding on the conformational stability of the diimine derivatives and (c) provide an entry to the potential use of these compounds in enantioselective recognition and catalysis. Since Clough and Roberts estimated the energy barrier to *syn/anti*-diastereomerization of 1,8-bis(2-methylphenyl)naphthalene, **4a**, as approximately 100 kJ/mol, (Clough, R. L.; Roberts, J. D. *J. Am. Chem. Soc.* **1976**, *98*, 1018-1020.) we expected that incorporation of methyl groups into the *ortho*-positions of **1** would produce conformational isomers that are stable to interconversion and separable at room temperature (Scheme 3). We therefore decided to prepare 1,8-bis(2'-methyl-4'-hydroxy-5'-formylphenyl)naphthalene, **5**, exhibiting moderate bulk adjacent to the chiral axes which should suffice to isolate and characterize the stereoisomers of this atropisomer and its diimine derivatives while racemization and diastereomerization reactions could be studied at elevated temperatures.



20 Scheme 3. Structures of **4a** exhibiting *anti*-parallel (*anti*-isomer) and parallel (*syn*-isomer) 2-methylphenyl moieties and of **5**.

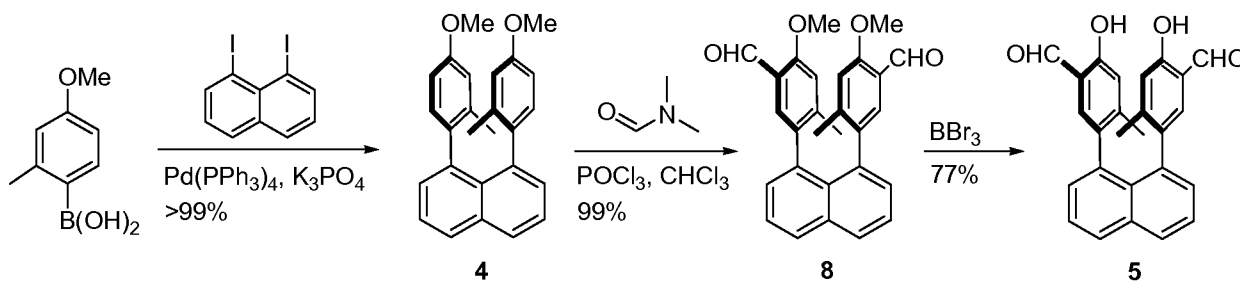
Based on our experience with the synthesis and atropisomerization of chiral biaryls and triaryls, (*E.g.*, Wolf, C.; Ghebramariam, B. T. *Tetrahedron: Asymm.* **2002**, *13*, 1153-1156.) and 1,8-diheteroarylnaphthalene-derived sensors, (*E.g.*, Mei, X.; Wolf, C. *Chem. Commun.* **2004**, 2078-2079.) we found that incorporation of steric bulk proximate to the aryl-aryl bonds in axially chiral 1,8-bisphenolnaphthalenes affords isolable enantiomers (Figure 1). We believe that this finding will have important implications to chiral ligand development for asymmetric catalysis, recognition and other fields. We now wish to report the first example of an application of a conformationally stable, axially chiral 1,8-bisphenolnaphthalene, such as **2** and **5**.

25



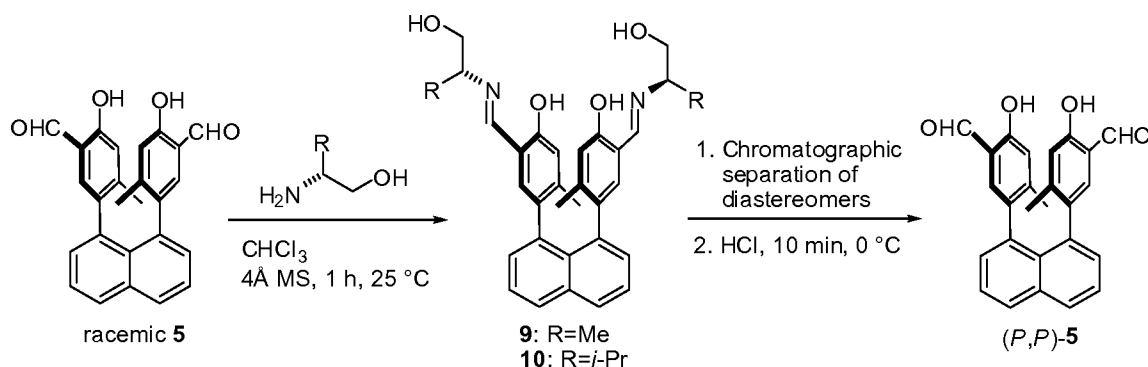
Structures of axially chiral 1,8-bisphenolnaphthalenes

We began the synthesis of **5** with Suzuki coupling of 1,8-diiodonaphthalene and commercially available 4-methoxy-2-methylphenylboronic acid (Scheme 4). Initially, we screened the effect of various catalysts, solvents, base and temperature to identify suitable reaction conditions for the construction of the sterically congested scaffold of **5**. We were pleased to find that 1,8-bis(2'-methyl-4'-methoxyphenyl)naphthalene, **4**, can be obtained in quantitative amounts using $\text{Pd}(\text{PPh}_3)_4$ as catalyst and K_3PO_4 as base in toluene. NMR analysis revealed that **4** was a 1:3 mixture of the *syn*- and *anti*-isomers. The Vilsmeier reaction with excess of phosphorous oxychloride and dimethyl formamide then furnished 1,8-bis(2'-methyl-4'-methoxy-5'-formylphenyl)naphthalene, **8**, with 99% yield in approximately the same diastereomeric ratio. Finally, deprotection with boron tribromide gave **5** having a *syn*- and *anti*-isomer ratio of 1:4 in 77% yield.

Scheme 4. Synthesis of 1,8-bis(2'-methyl-4'-hydroxy-5'-formylphenyl)naphthalene **5**.

The diastereomers of **8** were separated by column chromatography and the racemic *anti*-isomer was converted to *anti*-**5** and then to (*P,P,R,R*)-**9** and (*M,M,R,R*)-**9** by condensation with 2 equivalents of (*R*)-2-amino-1-propanol. The diimine formation proceeds with quantitative yields and is completed at room temperature within one hour. Chromatographic purification on silica gel then allowed isolation of the two diastereomeric products (Scheme 5). Heating of the diastereomeric mixture of **9** to establish thermodynamic equilibrium showed that the first eluted diimine corresponds to the more stable diastereomer, see below. The sense of amplification of asymmetric

induction observed with the diimines of **1** and CD and crystallographic analysis of **5** and **9** suggest that (*P,P,R,R*)-**9** is the thermodynamically favored atropisomer. Hydrolysis of (*P,P,R,R*)-**9** with aqueous HCl at 0 °C afforded enantiopure (*P,P*)-**5** in 85% yield, with no trace of the *syn*-diastereomer based on NMR analysis. The enantiopurity of **5** was confirmed by derivatization to the corresponding (*P,P,R,R*)-diimine with (*R*)-2-amino-1-propanol and NMR analysis did not show any signals of the diastereomeric (*M,M,R,R*)-isomer.



Scheme 5. Isolation of enantiopure **5**.

Based on our experience with stereolabile **1**, which spontaneously adopts a single conformation upon diimine formation with enantiopure amino alcohols and the kinetic analysis of **4a** by Clough and Roberts, we investigated the possibility to convert the atropisomeric mixture of **9** to a single isomer upon heating. Such an asymmetric transformation of the first kind would generate the thermodynamically favored diimine isomer and thus facilitate the formation of enantiopure **5** with a theoretical yield of 100% and without the need for an elaborate chromatographic separation of the equimolar mixture of (*M,M,R,R*)- and (*P,P,R,R*)-**9**. Several cases in which asymmetric transformation of the first kind was used to manipulate the diastereomeric ratio of axially chiral compounds have been reported. For example, Meyers *et al.* found that the stereochemical outcome of the diastereoselective oxazoline-mediated asymmetric Ullmann coupling of aryl bromides is significantly improved upon heating of the product mixture. (E.g., Nelson, T. D.; Meyers, A. I. *Tetrahedron Lett.* **1993**, 34, 3061-3062.) This transformation favors the formation of the desired (*P*)-atropisomer, a key intermediate for the total synthesis of permethylated tellimagrandin. (Nelson, T. D.; Meyers, A. I. *J. Org. Chem.* **1994**, 59, 2577-2580.) The same principle has been used for

5 the deracemization of *ortho*-dihydroxylated biaryl ligands VANOL and VAPOL (*E.g.*, Zhang, Y.; Yeung, S.-M.; Wu, H.; Heller, D. P.; Wu, C.; Wulff, W. D. *Org. Lett.* **2003**, 5, 1813-1816.) and for the preparation of the aglycon of vancomycin. (*E.g.*, Evans, D. A.; Wood, M. R.; Trotter, B. W.; Richardson, T. I.; Barrow, J. C.; Katz, J. L. *Angew. Chem., Int. Ed.* **1998**, 37, 2700-2704.)

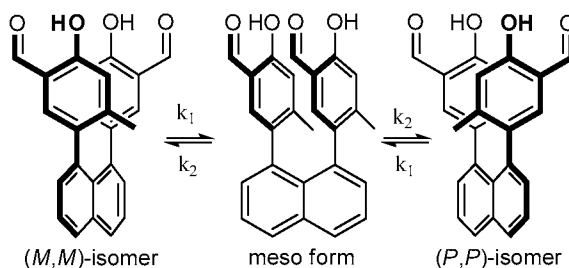
10 We realized that (*M,M,R,R*)- and (*P,P,R,R*)-**9**, formed from racemic **5** and (*R*)-2-amino-1-propanol at 25°C, showed distinct NMR spectra, for example two doublets at 1.26 and 1.40 ppm corresponding to the imino alcohol methyl groups (Figure 1). We therefore used NMR analysis to monitor the atropisomerization process. Upon heating to 60.0 °C, the signals of the thermodynamically less favored diastereomer decreased in
15 intensity and (*P,P,R,R*)-**9** with >98 %de was obtained after 14 hours (see Figure 1).

We previously reported that condensation of stereolabile **1** and (*R*)-2-amino-1-propanol exclusively generates the (*M,M,R,R*)-stereoisomer which is the thermodynamically favored conformer due to stabilization by intramolecular hydrogen bonding and concomitant minimization of steric repulsion. Accordingly, diimine
20 formation with (*S*)-2-amino-1-propanol gave the (*P,P,S,S*)-enantiomer, Scheme 2. These results are in perfect agreement with the asymmetric transformation of a mixture of (*M,M,R,R*)- and (*P,P,R,R*)-**9** towards the latter diastereomer. In analogy to the amplification of asymmetric induction observed with **1**, the central chirality of the imino alcohol moiety in **9** controls the chiral amplification and induces the same sense of axial
25 chirality. (Note that the (*M,M*)-scaffold in **1** corresponds to the (*P,P*)-conformation in **5** and **9** because the presence of the *ortho*-methyl groups in the latter results in a change in the CIP priorities.) Since the atropisomerization occurs with more than 99% de, it provides quantitative access to stereochemically pure **9** on the gram scale, which can then be hydrolyzed without concomitant isomerization to enantiopure **5**. Having
30 developed a convenient procedure producing (*P,P*)-**5**, we were able to prepare (*P,P,R,R*)-**9** and (*P,P,S,S*)-**9**, the thermodynamically less stable diastereomer, via condensation with either enantiomer of 2-amino-1-propanol. Analyzing the CD spectra of the enantiomeric (*M,M,R,R*)- and (*P,P,S,S*)-diimines of **1**, we previously speculated that the Cotton effects are predominantly controlled by the sense of axial chirality while the chiral centers in the
35 imino alcohol units were expected to have little or no effect on the chiroptical properties. Comparison of the CD spectra of (*P,P*)-**5**, (*P,P,R,R*)-**9** and (*P,P,S,S*)-**9**, all exhibiting the same sense of axial chirality, now clearly shows that this assumption is correct (Figure 2). The three atropisomers exhibit a pronounced positive Cotton effect, and the

5 incorporation of the diimino alcohol units results in a significant red shift. Importantly, the diastereomeric (*P,P*)-diimines of **9** show almost perfectly superimposable CD spectra, which underscores the overwhelming or possibly exclusive contribution of the relative orientation of the two cofacial salicylideneimine rings to the observed CD activity. (It is noteworthy that the CD amplitudes of the less stable (*P,P,S,S*)-**9** 10 diastereomer are slightly diminished compared to (*P,P,R,R*)-**9**. This is probably due to noticeable atropisomerization of (*P,P,S,S*)-**9** to its (*M,M,S,S*)-diastereomer, *i.e.* it is likely that the CD spectra obtained with (*P,P,S,S*)-**9** and (*P,P,R,R*)-**9** only differ because the former was not perfectly diastereomerically pure.)

Slow evaporation of a solution of enantiopure (*P,P*)-**5** in chloroform gave single 15 crystals suitable for X-ray studies (Figure 3). As expected, crystallographic analysis shows that the two salicylaldehyde rings reside in almost perfectly perpendicular orientation relative to the naphthalene backbone, exhibiting a C_2 -symmetric structure with a torsion angle of 5.32° . The splaying between the two phenyl rings was determined as 20.51° which results in a centroidal phenyl-to-phenyl separation of 3.47 Å. Based on 20 the enforced π -stacking of the proximate salicylaldehyde rings, the positive Cotton effect and the large CD amplitudes of (*P,P*)-**5** can be attributed to strong exciton coupling of the cofacial chromophores.

We then turned our attention to the kinetic analysis of **5** (Scheme 6). Interconversion of the stereoisomers of **5** requires one salicylaldehyde ring to rotate about the chiral 25 naphthyl-phenyl axis. Accordingly, the edge of the rotating ring points towards the adjacent phenyl moiety in the transition state. In general, this process can proceed via two T-shaped transition states having the methyl group of the rotating phenyl ring either directed towards or away from the other phenyl ring. (Wolf, C. (Ed.) Dynamic Stereochemistry of Chiral Compounds, RSC, Cambridge, 2008, pp. 89.) The latter 30 orientation is expected to afford significantly less steric hindrance and is therefore the favored interconversion pathway.



Scheme 6. Interconversion of the stereoisomers of **5**.

5

A solution of (*P,P*)-**5** in chloroform was stirred at 45.0 °C and small aliquots were taken at one hour intervals and diluted to 5.0×10^{-5} M for CD analysis. After 10 hours, the CD signals disappeared indicating complete racemization (Figure 4). The *syn/anti*-ratio of **5** at 45.0 °C in chloroform at equilibrium was determined by $^1\text{H-NMR}$ spectroscopy as 23.4:76.6. The observed ratio corresponds to a difference in Gibbs free energy of the *anti*- and *syn*-isomers, ΔG , of 1.3 kJ/mol according to the Boltzmann equation (1). (The factor 2 in equation 1 accounts for the two enantiomeric *anti*-isomers of **5**.)

$$2 N_{\text{syn}}/N_{\text{anti}} = \exp(-\Delta G^\circ/RT) \quad (1)$$

Figure 5 shows the decrease of the mole fraction of (*P,P*)-**5** as a function of time. The mathematical solution for the kinetics of consecutive, first-order, reversible reactions involving 3 species such as the *syn/anti*-interconversion of **5** has been reported by Vriens. (Vriens, G. N. *Ind. Eng. Chem.* **1954**, 669-671.) Curve fit analysis using equation 2 allowed determination of the rate constant for the *anti*- to *syn*-isomerization, k_1 .

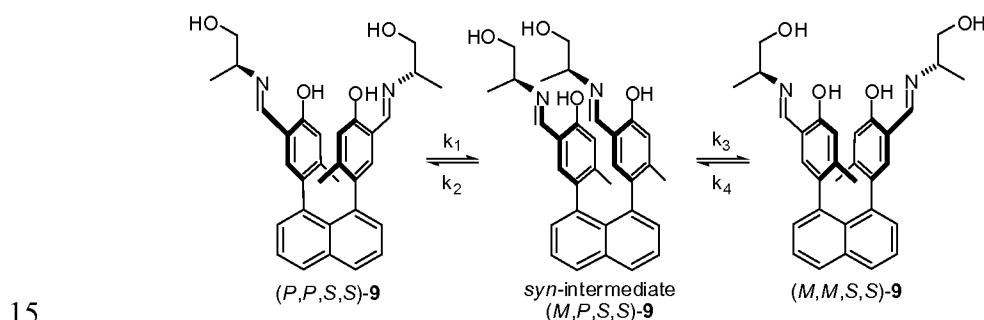
$$x = C_1 e^{D_1 k_1 t} + C_2 e^{D_2 k_1 t} + \frac{\alpha}{K_1 K_2 E_2} \quad (2)$$

k_1 = rate constant of the *anti*- to *syn*-interconversion, K_1 = equilibrium constant for the formation of the *syn*-isomer, K_2 = equilibrium constant for the formation of either *anti*-isomer, α = ratio of forward rate constants (k_2/k_1) for the consecutive, reversible, first-order reactions, k_2 = rate constant for *syn*- to *anti*-interconversion, C_1 , C_2 , D_1 , D_2 , E_2 are constants.

Having determined the *syn/anti*-ratio and thus the equilibrium constant for the isomerization of **5**, we were able to determine the rate constants for the reversible interconversion steps, k_1 and k_2 , as $6.308 \times 10^{-5} \text{ s}^{-1}$ for the *anti*→*syn*- and as $1.038 \times 10^{-4} \text{ s}^{-1}$ for the *syn*→*anti*-interconversion, respectively. As expected, the observed isomerizations proved to obey first-order kinetics. Using the Eyring equation, the Gibbs

- 5 activation energy, $\Delta G^\ddagger$, for the atropisomerization of **5** was calculated as 103.7 (102.4) kJ/mol for the conversion of the *anti*-(*syn*-) to the *syn*-(*anti*-) isomer (see SI).

The analysis of the atropisomerization of **9** is more complicated and involves four different rate constants (Scheme 7). Because CD analysis does not provide quantitative
 10 information about individual diastereomer concentrations we used NMR spectroscopy to monitor the conversion of (*P,P,S,S*)-**9**, which was prepared by condensation of (*P,P*)-**5** with 2 equivalents of (*S*)-2-amino-1-propanol, to the thermodynamically stable atropisomer (*M,M,S,S*)-**9** via the intermediate (*M,P,S,S*)-isomer.



Scheme 7. Interconversion of the atropisomers of **9**.

A solution of (*P,P,S,S*)-**9** in deuterated chloroform was heated to 58.0 °C and the isomerization was studied by integration of the benzylic ^1H NMR signals of the three
 20 diastereomers (see SI for details). Equilibrium was reached after 2 days, and the atropisomeric ratio was determined as 94.4:3.9:1.7 [(*M,M,S,S*):(*P,P,S,S*):(*M,P,S,S*)] (Figure 6). Accordingly, the thermodynamically favored (*M,M,S,S*)-atropisomer is more stable than the *syn*-intermediate by 11.2 kJ/mol while conversion of the latter to (*P,P,S,S*)-**9** is driven by only 2.4 kJ/mol. Comparison of the relative amounts of the two
 25 *anti*-isomers of **9** reveals a difference in Gibbs free energy of 8.8 kJ/mol. Crystallographic analysis of *syn*-**9** and (*P,P,R,R*)-**10** shows that these results can be explained by selective intramolecular hydrogen bonding and concomitant optimization of steric repulsion in (*M,M,S,S*)-**9**, *vide infra*. Following Vriens' mathematical treatment for two consecutive reversible reactions and curve fitting we then determined the individual
 30 rotational energy barriers (see SI for details). The interconversion of (*P,P,S,S*)-**9** to the *syn*-diastereomer has a Gibbs activation energy, $\Delta G^\ddagger_{(P,P,S,S)\text{-}\mathbf{9} \rightarrow (M,P,S,S)\text{-}\mathbf{9}}$, of 108.7 kJ/mol. The intermediate *syn*-isomer undergoes diastereomerization to the two *anti*-conformers

5 and the corresponding activation energies were calculated as 106.3 kJ/mol ($\Delta G^\ddagger_{(M,P,S,S)-9 \rightarrow (P,P,S,S)-9}$) and 104.5 kJ/mol ($\Delta G^\ddagger_{(M,P,S,S)-9 \rightarrow (M,M,S,S)-9}$). As expected from the asymmetric transformation experiments discussed above, the energy barrier for the conversion of (*M,M,S,S*)-**9** to the *syn*-isomer, $\Delta G^\ddagger_{(M,M,S,S)-9 \rightarrow (M,P,S,S)-9}$, is significantly higher and was determined as 115.7 kJ/mol. To confirm these data, we analyzed the initial decay of (*P,P,S,S*)-**9** at low conversion (less than 2% completion), which can be approximately
10 treated as an irreversible first-order reaction (see Experimental Section). We thus obtained a rotational energy barrier, $\Delta G^\ddagger_{(P,P,S,S)-9 \rightarrow (M,P,S,S)-9}$, of 109.2 kJ/mol, which is in very good agreement with the value determined by curve fitting.

Attempts to grow a single crystal of (*P,P,R,R*)- or (*M,M,S,S*)-**9** for crystallographic
15 analysis were unsuccessful. But we were able to obtain a crystal structure of the *syn*-isomer (Figure 7). This atropisomer has a torsion angle of 18.33° and the splaying between the two phenyl rings is 21.09° corresponding to a centroidal phenyl-to-phenyl separation of 3.52 Å. The steric repulsion between the two salicylideneimine rings explains the low relative stability compared to the (*P,P,R,R*)- or the (*M,M,S,S*)-isomer.

20 To better understand the overwhelming thermodynamic stability of the (*P,P,R,R*)- and the (*M,M,S,S*)-configuration, we decided to prepare the corresponding diimine using (*R*)-2-amino-3-methyl-1-butanol (see SI for details on the synthesis, CD analysis etc). Fortunately, a crystal of (*P,P,R,R*)-**10** was obtained by crystallization from a hexane solution (Figure 8). Crystallographic analysis revealed a torsion angle of 18.17° and
25 splaying between the two phenyl rings was calculated as only 13.46° resulting in a centroidal phenyl-to-phenyl separation of 3.33 Å. The significantly reduced splaying compared to *syn*-**9** is quite remarkable and results from reduced steric repulsion between the imino alcohol units and additional hydrogen bonding between the alcohol groups and the phenol units in the opposite salicylideneimine ring. The arrangement of the two
30 salicylideneimines in (*P,P,R,R*)-**10** allows formation of a total of four intramolecular hydrogen bonds (the phenol groups also undergo hydrogen bonding with the adjacent imines) while the steric bulk of the imino alcohol moieties is placed outside of this ring structure in the least crowded positions. The C=N $\cdots$ HOC_{phenyl} and the C_{aliph}OH $\cdots$ OC_{phenyl} hydrogen bond lengths are 1.861 and 2.178 Å, respectively.

35 In analogy to the results obtained with **9**, we observed quantitative asymmetric transformation of the first kind with **10**, which can be used to either prepare pure (*M,M,S,S*)- or (*P,P,R,R*)-atropisomers of these diimines. The kinetic and thermodynamic

5 analyses of the unidirectional atropisomerization process of **9** and **10** discussed above and in the SI are in perfect agreement with the crystallographic data showing distinct stabilization of the (*P,P,R,R*)-isomer due to intramolecular hydrogen bonding and minimized steric repulsion between the imine moieties.

We also found that (*P,P*)-**5** can be used for the kinetic resolution of the enantiomers
10 of 2-amino-1-propanol. Diimine formation of the enantiopure dialdehyde **5** and 4 equivalents of the racemic substrate at room temperature followed by extraction after 5 hours allowed recovery of the remaining amino alcohol in 47% yield. Chiral HPLC analysis showed that the (*S*)-amino alcohol was enriched to 54% ee, which proves the expected favored formation of (*P,P,R,R*)-**9** (see Experimental Section).

15 In conclusion, we have synthesized the first examples of axially chiral 1,8-bisphenolnaphthalenes that are stable to racemization at room temperature. The incorporation of two ortho-substituted phenol moieties into a rigid C₂-symmetric scaffold that is reminiscent of the successful BINOL motif has been of general interest due to potential applications in asymmetric catalysis and enantioselective sensing for a long
20 time. Stereochemical analysis showed that the *anti*-stereoisomers of 1,8-bis(2'-methyl-4'-hydroxy-5'-formylphenyl)naphthalene, **5**, and its diimine analogues **9** and **10** are significantly more stable than the corresponding *syn*-isomer. Slow *syn/anti*-interconversion, obeying reversible first-order kinetics, of **5**, **9** and **10** occurs at elevated temperatures and this provides a convenient entry towards enantiopure bisphenols that
25 can be prepared on the gram scale via asymmetric transformation of the first kind. Spectroscopic NMR and CD analyses supported by crystallography of *syn*- and *anti*-1,8-bisphenolnaphthalenes showed that the incorporation of two imino alcohol units into the triaryl scaffold controls the outcome of the unidirectional atropisomerization with literally perfect stereocontrol, resulting from a unique intramolecular hydrogen bonding
30 motif and concomitant reduction of steric repulsion. While the stereochemical bias of the atropisomers studied originates from the central chirality of the incorporated amino alcohol units, the chiroptical properties are solely determined by the sense of axial chirality. Chiral recognition studies using 1,8-bisphenolnaphthalenes as UV, CD and fluorescence sensors are currently underway in our laboratories.

35 The invention will now be further described by way of the following non-limiting examples.

EXAMPLES

5 A. Experimental Procedures for synthesizing 5, 9 and 10

All reagents and solvents were used without further purification. Reactions were carried out under nitrogen atmosphere and under anhydrous conditions. 1,8-Diiodonaphthalene was prepared from 1,8-diaminonaphthalene as described in the literature. (House, H. O.; Koepsella, D. G.; Campbel, W. J. J. Org. Chem. 1972, 37, 1003-1011.) Products were purified by flash chromatography on SiO₂ (particle size 0.032-0.063 mm). NMR spectra were obtained at 400 MHz (¹H NMR) and 100 MHz (¹³C NMR) using CDCl₃ as solvent. Chemical shifts are reported in ppm relative to TMS. For CD analysis, samples were diluted to 5.0 x 10⁻⁵ M with anhydrous chloroform and the instrument was purged with nitrogen for 20 minutes. Spectra were collected between 15 245 and 540 nm at 25.0 °C with a standard sensitivity of 100 mdeg, a data pitch of 0.5 nm, a band width of 1 nm, a scanning speed of 500 nm s⁻¹ and a response of 0.5 s using a quartz cuvette (1 cm path length). The data were adjusted by baseline correction and binomial smoothing.

1,8-Bis(2'-methyl-4'-methoxyphenyl)naphthalene (4)

20 A solution of 1,8-diiodonaphthalene, (1.70 g, 4.5 mmol), 2-methyl-4-methoxyphenylboronic acid, **2**, (2.20 g, 13.4 mmol), Pd(PPh₃)₄ (0.78 g, 0.67 mmol) and K₃PO₄, (4.30 g, 20.1 mmol) in 50 mL toluene was stirred at 100 °C for 18 hours. The resulting mixture was allowed to come to room temperature, quenched with water and extracted with CH₂Cl₂. The combined organic layers were dried over MgSO₄ and 25 concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂:hexanes 2:3) afforded 1.65 g (4.5 mmol, >99%) of off-white crystals containing *syn*- and *anti*-isomers of **7** in a ratio of 1:3.

¹H NMR: δ = 1.76 (s, 4.6H), 1.83 (s, 1.4H), 3.69 (s, 4.4H), 3.71 (s, 1.4H), 6.28-6.39 (m, 4H), 6.66 (d, *J* = 8.2 Hz, 0.5H), 6.87 (d, *J* = 8.2 Hz, 1.5H), 7.16 (d, *J* = 6.8 Hz, 2H), 7.46 (dd, *J* = 7.2, 7.8 Hz, 2H), 7.89 (d, *J* = 8.0, 2H). ¹³C NMR: δ = 20.9, 21.0, 55.1, 55.2, 109.9, 110.3, 114.3, 114.4, 124.8, 125.0, 128.5, 129.0, 130.2, 130.4, 131.0, 131.6, 132.3, 134.8, 134.9, 135.2, 135.4, 136.5, 136.9, 139.5, 157.6, 158.1. Anal. Calcd. for C₂₆H₂₄O₂: C, 84.75; H, 6.57. Found: C, 84.74; H, 6.61.

35 1,8-Bis(2'-methyl-4'-methoxy-5'-formylphenyl)naphthalene (8)

Phosphorous oxychloride (2.9 mL, 31.2 mmol) and dimethyl formamide (2.4 mL, 31.2 mmol) were stirred in 10 mL of chloroform at room temperature for one hour. Then, **7** (0.60 g, 1.6 mmol) was added and the mixture was refluxed at 90 °C for 48 hours. It

5 was then cooled to 0 °C, carefully quenched with water and extracted with CH₂Cl₂. The combined organic layers were dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂:EtOAc 25:1) afforded 0.69 g (1.6 mmol, 99%) of *syn/anti*-**8** as a white powder. The diastereomers of **8** can be separated using
10 flash chromatography with gradient elution starting with dichloromethane to collect the *anti*-isomer (73%), then increasing the polarity to DCM:EtOAc 15:1 to recover the *syn*-diastereomer (27%).

¹H NMR *anti*-**8**: δ = 1.81 (s, 6H), 3.80 (s, 6H), 6.35 (s, 2H), 7.15 (d, *J* = 7.0 Hz, 2H), 7.38 (s, 2H), 7.50 (dd, *J* = 7.0, 8.2 Hz, 2H), 7.94 (d, *J* = 8.2, 2H), 10.28 (s, 2H). ¹³C NMR: δ = 21.5, 55.5, 112.1, 121.6, 125.2, 127.8, 129.2, 130.1, 130.2, 134.8, 135.4,
15 137.4, 146.2, 160.3, 189.0. ¹H NMR *syn*-**8**: δ = 1.90 (s, 6H), 3.85 (s, 6H), 6.42 (s, 2H), 7.14 (d, *J* = 8.2 Hz, 2H), 7.15 (s, 2H), 7.49 (dd, *J* = 7.0, 8.2 Hz, 2H), 7.94 (d, *J* = 8.2, 2H), 10.17 (s, 2H). ¹³C NMR: δ = 21.6, 55.6, 112.0, 122.0, 125.0, 129.0, 130.3, 131.3, 131.8, 135.2, 137.5, 144.1, 159.5, 188.7. Anal. Calcd. for C₂₈H₂₄O₄: C, 79.22; H, 5.70. Found: C, 78.99; H, 5.72.

20

1,8-Bis(2'-methyl-4'-hydroxy-5'-formylphenyl)naphthalene (**5**)

To a solution of 1,8-bis(2'-methyl-4'-methoxy-5'-formylphenyl)naphthalene, **8**, (0.78 g, 1.9 mmol) in 35 mL of anhydrous CH₂Cl₂ at 0 °C, BBr₃ (11.8 mL, 11.8 mmol) was added dropwise and the mixture was stirred for six hours. The reaction was carefully
25 quenched with isopropyl alcohol followed by addition of water, and extracted with CH₂Cl₂. The combined organic layers were dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂:hexanes 20:1) afforded 0.58 g of **5** (1.5 mmol, 77%) as a white solid. The *anti/syn*-ratio was determined as 20:1 by ¹H-NMR spectroscopy. Enantiopure *anti*-**5** was obtained via formation of **9** or
30 **10** as described below and hydrolysis or by asymmetric transformation of the first kind and subsequent hydrolysis (see SI).

¹H NMR (*P,P*)-**5**: δ = 1.85 (s, 6H), 6.40 (s, 2H), 7.08 (s, 2H), 7.20 (d, *J* = 7.1 Hz, 2H), 7.54 (dd, *J* = 7.0 Hz, 7.1 Hz, 2H), 7.99 (d, *J* = 7.0 Hz, 2H), 9.66 (s, 2H), 10.73 (s, 2H). ¹³C NMR: δ = 21.7, 117.5, 118.0, 125.3, 125.4, 129.4, 129.5, 130.6, 132.9, 135.0,
35 136.9, 146.5, 160.5, 195.1. Anal. Calcd. for C₂₆H₂₀O₄: C, 78.77; H, 5.09. Found: C, 78.69; H, 5.42.

Diimine (**9**)

5 To racemic **5** (67.0 mg, 0.17 mmol) dissolved in 8 mL of CHCl₃ over molecular sieves (4 Å, beads, 8-12 mesh), 2 equivalents of (*R*)-2-amino-1-propanol (25.4 mg, 0.34 mmol) were added and the mixture was allowed to stir for one hour at room temperature. The mixture was then extracted with water, dried over MgSO₄ and concentrated in vacuo. Chromatographic purification using EtOAc:EtOH (99.5:0.5) as mobile phase
10 allowed the isolation of the two diastereomeric products (*P,P,R,R*)-**9** and (*M,M,R,R*)-**9** as yellow solids in quantitative yield.

¹H NMR (*P,P,R,R*)-**9**: δ = 1.26 (d, *J* = 6.4 Hz, 6H), 1.67 (s, 6H), 3.43 (t, *J* = 6.4 Hz, 2H), 3.62 (m, 2H), 3.72 (dd, *J* = 2.0 Hz, 12.0 Hz, 2H), 6.50 (s, 2H), 6.56 (s, 2H), 7.16 (d, *J* = 7.0 Hz, 2H), 7.49 (dd, *J* = 7.0 Hz, 8.0 Hz, 2H), 7.95 (m, 4H). ¹³C NMR: δ =
15 17.8, 20.8, 66.8, 67.2, 114.9, 117.8, 125.1, 129.0, 129.8, 130.8, 131.2, 133.4, 134.8, 138.0, 141.9, 161.8, 164.3. Anal. Calcd. for C₃₂H₃₄N₂O₄: C, 75.27; H, 6.71; N, 5.49. Found: C, 75.05; H, 6.98; N, 5.14.

Hydrolysis of (*P,P,R,R*)-**9** to (*P,P*)-**5**:

20 Pure (*P,P,R,R*)-**9** (43.2 mg, 0.08 mmol) was dissolved in 5 mL of 1 M NaOH, and 1 mL of aqueous HCl (12.1 M) was added dropwise at 0 °C. The mixture was allowed to stir for 10 minutes. The resulting suspension was extracted with CH₂Cl₂ and the combined organic layers were dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂) afforded enantiopure (*P,P*)-**5**
25 (28.5 mg, 0.07 mmol) in 85% yield, with no sign of the *syn*-diastereomer based on NMR analysis. The enantiopurity of **5** was confirmed by derivatization to the corresponding diimine using (*R*)-2-amino-1-propanol. The NMR spectrum of the condensation product showed the presence of a single diastereomer.

30

Diimine (*P,P,R,R*)-**10**

To racemic **5** (200 mg, 0.51 mmol) dissolved in 15 mL of CHCl₃ over molecular sieves (4 Å, beads, 8-12 mesh), 2 equivalents of (*R*)-2-amino-3-methyl-1-butanol (104 mg, 1.02 mmol) were added and the mixture was allowed to stir at 70 °C for 16 h. Upon
35 completion of the asymmetric transformation, the mixture was cooled to room temperature, extracted with water, dried over MgSO₄, and concentrated in vacuo. Chromatographic purification using CH₂Cl₂:EtOAc (1:1) as mobile phase gave (*P,P,R,R*)-**10** (285 mg, 0.50 mmol) in 99.8% yield.

5 ^1H NMR (*P,P,R,R*)-**10**: δ = 0.96 (d, J = 6.8 Hz, 12H), 1.64 (s, 6H), 1.91 (m, 2H), 3.05 (m, 2H), 3.68 (t, J = 10.5 Hz, 2H), 3.87 (m, 2H), 6.50 (s, 2H), 6.59 (s, 2H), 7.16 (d, J = 7.0 Hz, 2H), 7.50 (dd, J = 7.0 Hz, 8.0 Hz, 2H), 7.94 (m, 4H). ^{13}C NMR: δ = 18.6, 20.0, 20.8, 29.9, 64.2, 114.7, 118.1, 125.2, 128.9, 129.7, 130.9, 131.1, 133.2, 134.7, 138.1, 142.3, 163.0, 165.0. Anal. Calcd. for $\text{C}_{36}\text{H}_{42}\text{N}_2\text{O}_4$: C, 76.29; H, 7.47; N, 4.94. Found: C, 76.18; H, 7.28; N, 4.87.

The hydrolysis of (*P,P,R,R*)-**10** to (*P,P*)-**5** was conducted as described above with (*P,P,R,R*)-**9** and gave enantiopure (*P,P*)-**5** in 80% yield.

1. Deracemization of **5** via Asymmetric Transformation of the First Kind with **9**

Racemic **5** (10.14 mg, 0.026 mmol) was dissolved in 1.0 mL of CDCl_3 . Molecular sieves (4 Å, beads, 8-12 mesh) were added and the mixture was stirred with 2 equivalents of (*R*)-2-amino-1-propanol (3.84 mg, 0.051 mmol) for one hour at 25°C. After the diimine formation was complete, ^1H NMR analysis indicated the presence of two diastereomers – evidenced for example by the two doublets at 1.26 and 1.40 ppm. (Figure 1). Upon heating to 58.0 °C, the signals of the thermodynamically less favored diastereomer decreased in intensity. In agreement with the amplification of asymmetric induction observed upon diimine formation of 1,8-bis(3'-formyl-4'-hydroxyphenyl)naphthalene with (*R*)-2-amino-1-propanol, it is assumed that the central chirality of amino alcohol controls the stereoselective outcome of this atropisomerization process. The mixture is almost entirely converted to the more stable diastereomer after 14 hours (Figure 1).

Pure (*P,P,R,R*)-**9** (43.2 mg, 0.08 mmol) was dissolved in 5 mL of 1 M NaOH, and 1 mL of aqueous HCl (12.1 M) was added dropwise at 0 °C. The mixture was allowed to stir for 10 minutes. The resulting suspension was extracted with CH_2Cl_2 and the combined organic layers were dried over MgSO_4 and concentrated in vacuo. Purification by flash chromatography on silica gel (CH_2Cl_2) afforded enantiopure (*P,P*)-**5** (28.5 mg, 0.07 mmol) in 85% yield, with no sign of the *syn*-diastereomer based on NMR analysis. The enantiopurity of **5** was confirmed by derivatization to the corresponding diimine using (*R*)-2-amino-1-propanol. The NMR spectrum of the condensation product showed the presence of a single diastereomer.

2. Deracemization of **5** via Asymmetric Transformation of the First Kind with **10**

5 The diimine (*P,P,R,R*)-**10** was prepared in quantitative yields from racemic **5** and (*R*)-2-amino-3-methyl-1-butanol using the same asymmetric transformation protocol as described above for **9**. The hydrolysis of (*P,P,R,R*)-**10** to (*P,P*)-**5** was conducted as described above with (*P,P,R,R*)-**9** and gave enantiopure (*P,P*)-**5** in 80% yield. The CD spectra of (*P,P,R,R*)-**10** and (*P,P*)-**5** are shown in Figure 9.

10 The unidirectional atropisomerization of (*P,P,S,S*)-**10** to (*M,M,S,S*)-**10** was monitored by CD spectroscopy and is in perfect agreement with the results obtained with **9**. Enantiopure (*P,P*)-**5** (10.14 mg, 0.026 mmol) was dissolved in 1.0 mL of CDCl₃. Molecular sieves (4 Å, beads, 8-12 mesh) were added and the mixture was stirred with 2 equivalents of (*S*)-2-amino-3-methyl-1-butanol (5.52 mg, 0.051 mmol) for one hour at
15 25.0 °C. The mixture was then heated to 50.0 °C. Aliquots were taken at one hour intervals and diluted to 5.0 x 10⁻⁵ M with anhydrous CHCl₃ for CD analysis. After 24 hours, the CD signal indicated almost complete diastereomerization (Figure 10).

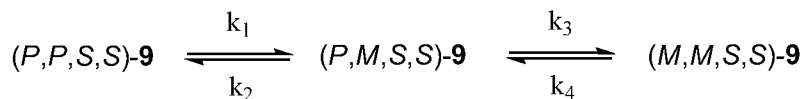
3. Determination of the Initial Rate Constant for the Isomerization of (*P,P,S,S*)-**9**

20 Enantiopure (*P,P*)-**5** (10.14 mg, 0.026 mmol) was dissolved in 1.0 mL of CDCl₃. Molecular sieves (4 Å, beads, 8-12 mesh) were added and the mixture was stirred with 2 equivalents of (*S*)-2-amino-1-propanol (3.84 mg, 0.051 mmol) for one hour at 25 °C. The mixture was then heated to 58.0 °C and ¹H-NMR spectra were collected at short intervals to follow the decay of (*P,P,S,S*)-**9** (same signals as (*M,M,R,R*)-**9**, Figure 1) before the
25 appearance of any (*M,M,S,S*)-**9** (same signals as (*P,P,R,R*)-**9**, Figure 1). By plotting the mol fraction of (*P,P,S,S*)-**9** versus time, the initial rate of the reaction can be obtained from the slope of the fitted line (Figure 11). Curve fitting to $y = A_1x + B$ was performed using OriginPro 8.1, with $A_1 = -0.00245$ and $B = 0.8634$, $A_2 = 0.351$, with $R^2 = 0.9929$.

The initial rate constant denoted k_1 was determined as $4.083 \times 10^{-5} \text{ s}^{-1}$, $\Delta G^\ddagger_{(P,P,S,S)\text{-}9 \rightarrow (P,M,S,S)\text{-}9} = 109.2 \text{ kJ/mol}$.

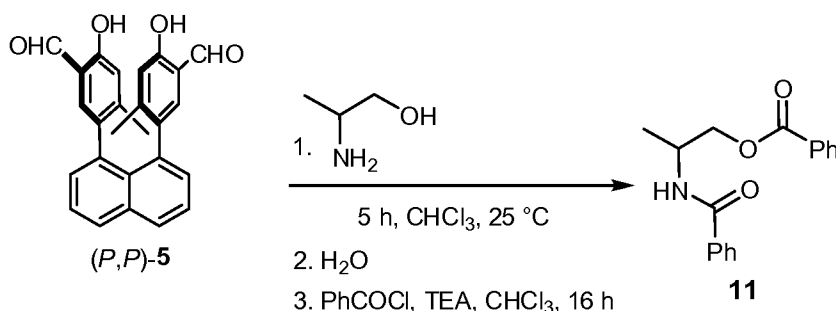
4. Determination of the Rate Constants for the Isomerization of **9**²⁰

As described above, enantiopure (*P,P*)-**5** (10.12 mg, 0.026 mmol) was dissolved in 1.0 mL of CDCl₃, treated with 2 equivalents of (*S*)-2-amino-1-propanol (3.84 mg, 0.051 mmol) for one hour at 25 °C, and then heated to 58.0 °C. ¹H-NMR spectra were collected
35 at different intervals to monitor the change in the intensity of the resolved methyl signals of (*P,P,S,S*)-**9** (1.69 ppm), (*M,P,S,S*)-**9** (1.81 ppm) and (*M,M,S,S*)-**9** (1.66 ppm) until equilibrium was reached.



The relative amounts of the three stereoisomers at equilibrium were determined as 94.4:3.9:1.7 [(*M,M,S,S*):(*P,P,S,S*):(*M,P,S,S*)] in CDCl₃ at 58.0 °C. The individual rate constants were then calculated as described in the SI.

5. Kinetic Resolution



Scheme 8. Kinetic resolution of 2-amino-1-propanol using (*P,P*)-**5**.

Enantiopure (*P,P*)-**5** (15.89 mg, 0.040 mmol) was dissolved in 1.0 mL of anhydrous CHCl₃. Molecular sieves (4 Å, beads, 8-12 mesh) were added and the mixture was stirred with 4 equivalents of racemic 2-amino-1-propanol (12.04 mg, 0.160 mmol) for 5 hours at 25 °C. The mixture was then extracted with water, and the aqueous layer was freeze-dried to recover the unreacted amino alcohol. The crude material (6.02 mg, 0.080 mmol) was dissolved in 1.5 mL of chloroform and treated with benzoyl chloride (117.8 mg, 0.801 mmol) in the presence of triethylamine (162.2 mg, 1.60 mmol). The mixture was allowed to stir for 16 hours, then quenched with water, and extracted with CH₂Cl₂. The combined organic layers were dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂:EtOAc 9:1) afforded **11** (22.0 mg, 0.074 mmol) as a colorless oil in 93%. The ee was determined as 54% by HPLC on Chiralpak AD using hexanes:ethanol (90:10) as mobile phase. As expected, formation of (*P,P,R,R*)-**9** is favored and comparison of the HPLC chromatogram with separately prepared enantiopure samples of **11** proved that 77% of the unreacted amino alcohol had (*S*)-configuration.

6. Crystallization and X-Ray Diffraction

Single crystal X-ray analysis was performed at 100 K using a Siemens platform diffractometer with graphite monochromated Mo-Kα radiation (λ = 0.71073 Å). Data were integrated and corrected using the Apex 2 program. The structures were solved by

- 5 direct methods and refined with full-matrix least-square analysis using SHELX-97-2 software. Non-hydrogen atoms were refined with anisotropic displacement parameters.

A crystal of enantiopure (*P,P*)-**5** was obtained by slow evaporation of a solution of 5.0 mg of (*P,P*)-**5** in 3 mL of CHCl₃. Crystal structure data for (*P,P*)-**5**: Formula C₂₆H₂₀O₄, M= 396.43, crystal dimensions 0.15 x 0.10 x 0.05 mm, tetragonal, space group
 10 P4₃, *a* = 11.7955(17) Å, *b* = 11.7955 (17) Å, *c* = 28.4769(40) Å, *V* = 3962.10 Å³, *Z* = 1, $\rho_{\text{calcd}} = 1.3290 \text{ g cm}^{-3}$.

Selected distances and angles:

Phenyl-phenyl [Å] (centroids to centroids)	3.470
Splaying angle between salicylidenimine planes [°]	20.51
Torsion angle [°]	5.32 15

The slow evaporation of a solution of 5.0 mg of **5** and 2 equivalents of (*R*)-2-amino-1-propanol in 3 mL of CHCl₃ afforded single crystals of *syn*-**9**. Crystal structure data for *syn*-**9**: Formula C₃₂H₃₄N₂O₄, M=510.62, crystal dimensions 0.120 x 0.10 x 0.07 mm,
 20 monoclinic, space group P2₁/c, *a* = 22.3400(21) Å, *b* = 6.8469 (6) Å, *c* = 17.6753 (16) Å, $\beta = 100.956 (1)$, *V* = 2654.33 Å³, *Z* = 4, $\rho_{\text{calcd}} = 1.2726 \text{ g cm}^{-3}$.

Selected distances and angles:

Phenyl-phenyl [Å] (centroids to centroids)	3.518
Splaying angle between salicylidenimine planes [°]	21.09
Torsion angle [°]	18.33 25

A crystal of (*P,P,R,R*)-**10** was obtained by crystallization of 100.0 mg of (*P,P,R,R*)-**10** from 15 mL of hexanes. Crystal structure data for (*P,P,R,R*)-**10**: Formula C₃₂H₃₄N₂O₄, M=566.73, crystal dimensions 0.10 x 0.10 x 0.05 mm, orthorhombic, space group P2₁, *a*
 30 = 10.8469(44) Å, *b* = 21.3655 (86) Å, *c* = 13.6440 (55) Å, *V* = 3161.99 Å³, *Z* = 2, $\rho_{\text{calcd}} = 1.1903 \text{ g cm}^{-3}$.

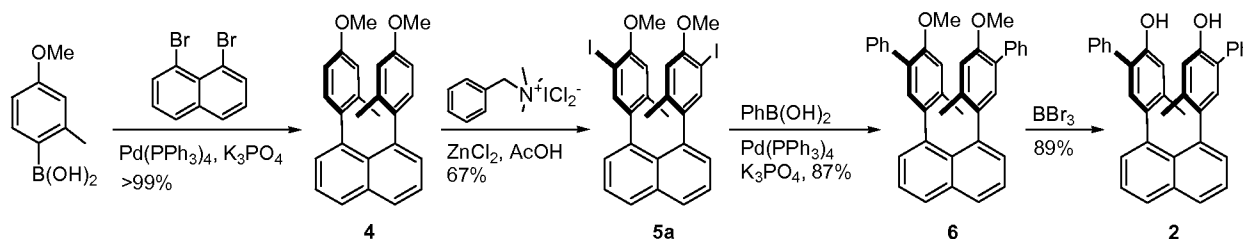
Selected distances and angles:

C=N $\cdots$ HOC _{phenyl} hydrogen bond length [Å]	1.861
C _{aliph} OH $\cdots$ OC _{phenyl} hydrogen bond length [Å]	2.178
Phenyl-phenyl [Å] (centroids to centroids)	3.326 35
Splaying angle between salicylidenimine planes [°]	13.46
Torsion angle [°]	18.17

5

B. Experimental Procedures for synthesizing 2

All reagents and solvents were used without further purification. All reactions were carried out under nitrogen atmosphere and anhydrous conditions. Products were purified by flash chromatography on SiO₂ (particle size 0.032-0.063 mm). NMR spectra were obtained at 400 MHz (¹H NMR) and 100 MHz (¹³C NMR) using CDCl₃ as solvent. Chemical shifts are reported in ppm relative to TMS.

**Synthesis of 1,8-bisphenolnaphthalene 2.**

15

1,8-Bis(2'-methyl-4'-methoxyphenyl)naphthalene (4)

A solution of 1,8-dibromonaphthalene, (1.29 g, 4.5 mmol), 2-methyl-4-methoxyphenylboronic acid, **3**, (2.24 g, 13.5 mmol), Pd(PPh₃)₄ (0.78 g, 0.68 mmol) and K₃PO₄, (4.30 g, 20.0 mmol) in 50 mL toluene was stirred at 110 °C for 18 hours. The resulting mixture was allowed to come to room temperature, quenched with water and extracted with CH₂Cl₂. The combined organic layers were dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂:hexanes 2:3) afforded 1.66 g (4.5 mmol, >99%) of off-white crystals containing *syn*- and *anti*-isomers of **4**.

¹H NMR: δ = 1.76 (s, 4.6H), 1.83 (s, 1.4H), 3.69 (s, 4.4H), 3.71 (s, 1.4H), 6.28-6.39 (m, 4H), 6.66 (d, *J* = 8.2 Hz, 0.5H), 6.87 (d, *J* = 8.2 Hz, 1.5H), 7.16 (d, *J* = 6.8 Hz, 2H), 7.46 (dd, *J* = 6.8 Hz, 8.0 Hz, 2H), 7.89 (d, *J* = 8.0 Hz, 2H). ¹³C NMR: δ = 20.9, 21.0, 55.1, 55.2, 109.9, 110.3, 114.3, 114.4, 124.8, 125.0, 128.5, 129.0, 130.2, 130.4, 131.0, 131.6, 132.3, 134.8, 134.9, 135.2, 135.4, 136.5, 136.9, 139.5, 157.6, 158.1. Anal. Calcd. for C₂₆H₂₄O₂: C, 84.75; H, 6.57. Found: C, 84.74; H, 6.61.

30

1,8-Bis(2'-methyl-4'-methoxy-5'-iodophenyl)naphthalene (5^a)

To a solution of **4** (0.18 g, 0.49 mmol) in 8 mL of acetic acid, benzyltrimethylammonium dichloroiodate (0.38 g, 1.08 mmol) and zinc dichloride (0.15

5 g, 1.08 mmol) dissolved in 8 mL of acetic acid were added dropwise over 30 minutes, and the mixture was stirred at 55 °C for 20 hours. It was then cooled to 0 °C, carefully quenched with water and extracted with CH₂Cl₂. The combined organic layers were washed with 1 M sodium thiosulfate, dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (hexanes:CH₂Cl₂ 3:1) afforded 0.20 g
10 (0.33 mmol, 67%) of **5a** as a white solid.

¹H NMR: δ = 1.86 (s, 6H), 3.83 (s, 6H), 6.41 (s, 2H), 7.15 (d, *J* = 6.9 Hz, 2H), 7.27 (s, 2H), 7.48 (dd, *J* = 6.9 Hz, 8.1 Hz, 2H), 7.92 (d, *J* = 8.1 Hz, 2H). ¹³C NMR: δ = 20.7, 56.2, 82.0, 111.2, 125.0, 129.0, 130.2, 134.9, 136.8, 137.5, 137.6, 138.3, 156.2. Anal. Calcd. for C₂₆H₂₂I₂O₂: C, 50.35; H, 3.58. Found: C, 50.15; H, 3.52.

15

1,8-Bis(6'-methoxy-4'-methylbiphenyl-3'-yl)naphthalene (6)

A solution of **5a** (0.20 g, 0.32 mmol), phenylboronic acid (0.12 g, 0.96 mmol), Pd(PPh₃)₄ (0.055 g, 0.048 mmol) and K₃PO₄ (0.31 g, 1.44 mmol) in 5 mL of toluene was stirred at 110 °C for 18 hours. The resulting mixture was allowed to come to room
20 temperature, quenched with water and extracted with CH₂Cl₂. The combined organic layers were dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂:hexanes 1:2) afforded 0.14 g (0.28 mmol, 87%) of **6** as a white solid.

¹H NMR: δ = 1.70 (s, 6H), 3.67 (s, 6H), 6.36 (s, 2H), 7.00 (s, 2H), 7.21-7.31 (m, 6H), 7.38 (dd, *J* = 7.5 Hz, 7.7 Hz, 4H), 7.45-7.52 (m, 6H), 7.92 (d, *J* = 8.2 Hz, 2H). ¹³C
25 NMR: δ = 20.7, 55.5, 125.1, 126.3, 126.5, 127.9, 128.6, 129.2, 130.2, 130.6, 130.9, 134.8, 135.3, 136.4, 138.6, 139.1, 154.7. Anal. Calcd. for C₃₈H₃₂O₂: C, 87.66; H, 6.19. Found: C, 87.42; H, 6.47.

1,8-Bis(6'-hydroxy-4'-methylbiphenyl-3'-yl)naphthalene (2)

To a solution of 1,8-bis(2'-methyl-4'-methoxy-5'-phenylphenyl)naphthalene, **6**, (0.14 g, 0.28 mmol) in 10 mL of anhydrous CH₂Cl₂ at 0 °C, BBr₃ (1.67 mL, 1.67 mmol) was added dropwise and the mixture was stirred for sixteen hours at room temperature. The reaction was carefully quenched with isopropyl alcohol followed by addition of
35 water, and extracted with CH₂Cl₂. The combined organic layers were dried over MgSO₄ and concentrated in vacuo. Purification by flash chromatography on silica gel (CH₂Cl₂) afforded 0.12 g of **2** (0.25 mmol, 89%) as a white solid.

5 ^1H NMR: δ = 1.66 (s, 6H), 5.01 (s, 2H), 6.45 (s, 2H), 6.89 (s, 2H), 7.25 (d, J = 7.0 Hz, 2H), 7.36 (dd, J = 6.8 Hz, 7.2 Hz, 2H), 7.42-7.51 (m, 10H), 7.91 (d, J = 8.1 Hz, 2H). ^{13}C NMR: δ = 20.3, 115.9, 124.2, 125.0, 127.5, 128.6, 128.8, 129.2, 129.9, 130.3, 130.9, 150.7. Anal. Calcd. for $\text{C}_{36}\text{H}_{28}\text{O}_2$: C, 87.78; H, 5.73. Found: C, 88.04; H, 5.79.

10 **1,8-Bis(6'-*N*-Boc-L-tryptophan-4'-methylbiphenyl-3'-yl)naphthalene (7)**

A solution of **6** (0.17 g, 0.35 mmol), *N*-Boc-L-tryptophan (0.24 g, 0.77 mmol), *N,N'*-dicyclohexylcarbodiimide (0.17 g, 0.81 mmol) and DMAP (0.05 g, 0.42 mmol) was stirred in 5 mL of dichloromethane for 16 hours at room temperature. The resulting suspension was dried and subjected to gradient flash chromatography on silica gel
15 (dichloromethane:ethyl acetate 70:1 to 60:1) to afford 0.18 g (0.17 mmol, 49%) and 0.16 g (0.15 mmol, 42%) of the first and second eluting diastereomers of **7** as white solids.

^1H NMR (first eluting): δ = 1.38 (s, 6H), 1.43 (s, 18H), 3.22 (m, 4H), 4.26 (m, 1H), 5.01 (m, 1H), 5.63 (s, 2H), 6.79 (s, 2H), 6.97-7.28 (m, 26H), 7.94 (d, J = 8.1 Hz, 2H), 9.33 (s, 2H). ^{13}C NMR: δ = 20.1, 28.3, 54.5, 80.0, 109.8, 111.3, 118.7, 119.5, 121.9,
20 122.0, 123.3, 124.9, 127.0, 127.7, 128.4, 128.7, 121.9, 123.3, 124.9, 127.0, 127.7, 128.7, 129.8, 131.2, 136.2, 137.3, 138.1, 140.5, 145.5, 155.2, 170.9. Anal. Calcd. for $\text{C}_{68}\text{H}_{64}\text{N}_4\text{O}_8$: C, 76.67; H, 6.06; N, 5.26. Found: C, 76.44; H, 6.16; N, 5.62.

^1H NMR (second eluting): δ = 1.40 (s, 18H), 1.70 (s, 6H), 3.12 (m, 4H), 4.82 (m, 1H), 5.02 (m, 1H), 6.65 (s, 2H), 6.88 (s, 2H), 7.11-7.31 (m, 18H), 7.51 (m, 2H), 7.92 (d,
25 J = 8.2 Hz, 2H), 8.41 (s, 2H). ^{13}C NMR: δ = 20.1, 28.3, 54.5, 80.0, 109.8, 111.1, 118.7, 119.5, 122.0, 122.9, 123.1, 124.9, 127.3, 127.9, 128.2, 128.8, 130.3, 131.5, 135.2, 136.0, 136.1, 137.2, 138.1, 140.51, 145.7, 155.46, 170.4. Anal. Calcd. for $\text{C}_{68}\text{H}_{64}\text{N}_4\text{O}_8$: C, 76.67; H, 6.06; N, 5.26. Found: C, 76.43; H, 6.27; N, 5.42.

30 **1,8-Bis(6'-hydroxy-4'-methylbiphenyl-3'-yl)naphthalene (2)**

A suspension of the first eluting diastereomer of **7** (0.04 g, 0.038 mmol) in 3.8 M KOH (10 mL, 4:1 ethanol:water) was stirred for 3 hours at room temperature. The resulting mixture was quenched with 0.5 mL of concentrated HCl at 0 °C. It was then extracted with dichloromethane, and the combined organic layers were washed with
35 brine, dried over MgSO_4 and concentrated in vacuo. The crude was then subjected to gradient flash chromatography on silica gel (dichloromethane) to afford 0.019 g (0.038 mmol, 99%) of enantiopure levorotatory-**2** as a white solid. The same procedure was used to obtain dextrorotatory-**2** by hydrolyzing the second eluting diastereomer of **7**.

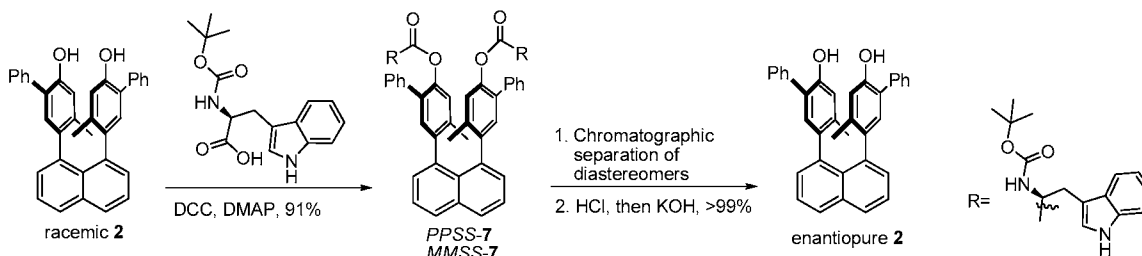
¹H NMR: δ = 1.66 (s, 6H), 5.01 (s, 2H), 6.45 (s, 2H), 6.89 (s, 2H), 7.25 (d, J = 7.0 Hz, 2H), 7.36 (dd, J = 6.8 Hz, 7.2 Hz, 2H), 7.42-7.51 (m, 10H), 7.91 (d, J = 8.1 Hz, 2H). ¹³C NMR: δ = 20.3, 115.9, 124.2, 125.0, 127.5, 128.6, 128.8, 129.2, 129.9, 130.3, 130.9, 150.7. Anal. Calcd. for C₃₆H₂₈O₂: C, 87.78; H, 5.73. Found: C, 88.04; H, 5.79.

1. Resolution of 2

We found that the enantiomers and the meso *syn*-isomer of **2** can be resolved by HPLC on Chiralpak AD or other means (see SI). Preparative isolation of the levorotatory enantiomer enabled us to monitor the racemization kinetics at 77.1 °C by chiral HPLC separation of cooled aliquots (Figure 12). The *syn/anti*-ratio of **2** at equilibrium was determined as 0.46 and the reversible first order reaction kinetics were simulated and analyzed according to Vriens. (Vriens, G. N. *Ind. Eng. Chem.* **1954**, 669-671.) The energy barrier to conversion of the *trans*-isomer to the *syn*-intermediate, $\Delta G^{\ddagger}_{anti \rightarrow syn}$, was determined as 110.0 kJ/mol. As expected, the Gibbs activation energy for the reversed process, $\Delta G^{\ddagger}_{syn \rightarrow anti}$, is slightly lower and was calculated as 107.7 kJ/mol (see SI).

The stereoisomers of **2** are separable on a CHIRALPAK AD column using hexanes:isopropyl alcohol (9:1) as mobile phase (Figure 13). The first eluting enantiomer elutes at 7.2 min, while the second elutes at 8.9 min. The *syn*-intermediate elutes last at 10.5 min. The absorption coefficient of the *syn*-diastereomer was determined as 1124216.92 unit area/ μ mol, while that for the *anti*-isomer was 1607421.75 unit area/ μ mol.

To avoid this tedious separation process, several chiral auxiliaries were examined to allow for the large scale separation of enantiomers of **2** through the formation of diastereomeric complexes. After screening several chiral auxiliaries, *N*-Boc-L-tryptophan was used to esterify racemic **2**, and was followed by the chromatographic separation of the resulting diastereomers (Scheme 2). Finally, the diastereomers were hydrolyzed to regenerate the free phenols in their enantiomerically pure forms.



Resolution of **2** through the formation of diastereomeric intermediates

5

2. UV, CD and Polarimetry

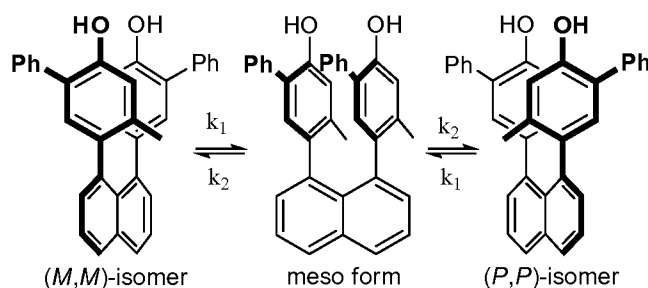
The CD instrument was purged with nitrogen for 20 minutes. Spectra were collected at room temperature between 270 and 390 nm with a standard sensitivity of 100 mdeg, a data pitch of 0.5 nm, a band width of 1 nm, a scanning speed of 500 nm s⁻¹ and a response of 0.5 s using a quartz cuvette (1 cm path length). Polarimetric measurements at 21.5 °C allowed for the determination of the specific rotation $[\alpha]_D$ as 354 for the first eluting enantiomer, and -355 for the second. The CD spectra of the enantiomers of **2** are shown in Figures 14. The effects on the CD spectra of the enantiomers of **2** upon treatment of sodium *ter*-butoxide is shown in Figure 15.

3. Kinetic Studies

3.1. Isomerization experiment

Several solutions of the levorotatory enantiomer of **1** (1.02×10^{-3} M, isopropyl alcohol) were prepared in separate vials and stirred at 77.1 °C (± 0.4 °C). The temperature was continuously monitored using a calibrated digital thermometer. At various times, the stereoisomeric composition was determined using a rapidly cooled vial. From this vial, 50 µL were diluted with 1 mL of hexanes:IPA (9:1), and then analyzed by HPLC on CHIRALPAK AD and UV detection at 240 nm. (see Figure 16)

3.2. Determination of the rate constants for the isomerization of **2**



25

$$x = C_1 e^{D_1 k_1 t} + C_2 e^{D_2 k_1 t} + \frac{\alpha}{K_1 K_2 E_2} \quad (1)$$

k_1 = rate constant for *anti*- to *syn*-interconversion

k_2 = rate constant for *syn*- to *anti*-interconversion

K_1 = equilibrium constant for the formation of the *syn*-isomer

30 K_2 = equilibrium constant for the formation of the *anti*-isomer

α = ratio of forward rate constants for the consecutive reactions

5

Figure 17 shows the decrease of the mol fraction of (-)-**2** as a function of time. The mathematical solution for the kinetics of consecutive, first-order, reversible reactions involving 3 species such as the *syn/anti*-interconversion of **2** has been reported by Vriens.⁷ Curve fit analysis using equation 2 allows for the determination of the rate constant for the *anti*- to *syn*-isomerization, k_1 .

$$x = C_1 e^{D_1 k_1 t} + C_2 e^{D_2 k_1 t} + \frac{\alpha}{K_1 K_2 E_2} \quad (2)$$

k_1 = rate constant of the *anti*- to *syn*-interconversion, K_1 = equilibrium constant for the formation of the *syn*-isomer, K_2 = equilibrium constant for the formation of either *anti*-isomer, α = ratio of forward rate constants (k_2/k_1) for the consecutive, reversible, first-order reactions, k_2 = rate constant for *syn*- to *anti*-interconversion, C_1 , C_2 , D_1 , D_2 , E_2 are constants. Curve fitting of the decay of (-)-**2** to $y = A1 \cdot \exp(-x/t1) + A2 \cdot \exp(-x/t2) + y0$ was performed using OriginPro 8.1, with $A1 = 0.40745$, $A2 = 0.11968$, $t1 = 77.31456$, $t2 = 8.33489$, $y0 = 0.40745$, $R^2 = 0.9999$.

The *syn/anti*-ratio of **2** at equilibrium was determined as 0.46. The mixture consists of 81.4% of racemic *anti*-**2** and 19.1% of *syn*-**2**.

$$\begin{aligned} E_1 &= 1 + \frac{1}{K_1} + \alpha + \frac{\alpha}{K_2} & E_2 &= \alpha \left(1 + \frac{1}{K_1 K_2} + \frac{1}{K_2} \right) = D_1 D_2 \\ D_1 &= \frac{-E_1 + \sqrt{E_1^2 - 4E_2}}{2} & D_2 &= \frac{-E_1 - \sqrt{E_1^2 - 4E_2}}{2} \\ C_1 &= \frac{-1 - D_2 + \frac{\alpha}{K_1 K_2 D_1}}{D_1 - D_2} & C_2 &= \frac{1 + D_1 - \frac{\alpha}{K_1 K_2 D_2}}{D_1 - D_2} \end{aligned}$$

$K_1 = 0.466$ and $K_2 = 1/K_1 = \alpha = 2.147$.

25

Using the above equations, the remaining constants can now be calculated.

$E_1 = 6.29$, $E_2 = 5.29$

$D_1 = -1.00$, $D_2 = -5.29$

$C_1 = 0.500$, $C_2 = 0.094$

30

$$5 \quad D_2k_1 = -1/t_2 = -0.1199, k_1 = 0.0003778 \text{ s}^{-1}$$

$$K^{\ddagger}_{anti \rightarrow syn} = k_1 h / k_b T = 5.18 \times 10^{-17}$$

$$\Delta G^{\ddagger}_{anti \rightarrow syn} = -RT \ln K^{\ddagger}_{anti \rightarrow syn} = 109.2 \text{ kJ.mol}^{-1}.$$

$$\text{Also, } D_1k_1 = -1/t_1 = -0.01293, k_1 = 0.00021557 \text{ s}^{-1}$$

$$K^{\ddagger}_{anti \rightarrow syn} = k_1 h / k_b T = 2.95 \times 10^{-17}$$

$$10 \quad \Delta G^{\ddagger}_{anti \rightarrow syn} = -RT \ln K^{\ddagger}_{anti \rightarrow syn} = 110.8 \text{ kJ.mol}^{-1}.$$

$$\text{Averaging of these two values gives } \Delta G^{\ddagger}_{anti \rightarrow syn} = 110.0 \text{ kJ.mol}^{-1}.$$

$$\text{Since } k_2 = k_1/K_1, k_2 = 0.0006373 \text{ s}^{-1}$$

$$K^{\ddagger}_{syn \rightarrow anti} = k_2 h / k_b T = 8.73 \times 10^{-17}$$

$$15 \quad \Delta G^{\ddagger}_{syn \rightarrow anti} = -RT \ln K^{\ddagger}_{syn \rightarrow anti} = 107.7 \text{ kJ.mol}^{-1}.$$

5 4. Crystallization and X-Ray Diffraction

Having determined the stability of **2** to rotation about the chiral axes at room temperature, we decided to investigate the three-dimensional structure of this atropisomer. A single crystal of *anti*-**2** was obtained by slow evaporation of a dichloromethane solution. Crystallographic analysis revealed that the two cofacial phenol rings are almost perfectly aligned due to a torsion angle of only 0.28°. The separation between the centroids of the two phenol rings, which are splayed by 19.48°, was determined as 3.5 Å, indicating substantial π -overlap (Figure 18).

Single crystal X-ray analysis was performed at 100 K using a Siemens platform diffractometer with graphite monochromated Mo-K α radiation ($\lambda = 0.71073$ Å). Data were integrated and corrected using the Apex 2 program. The structures were solved by direct methods and refined with full-matrix least-square analysis using SHELX-97-2 software. Non-hydrogen atoms were refined with anisotropic displacement parameters.

A crystal of **2** was obtained by slow evaporation of a solution of 10.0 mg of **2** in 5 mL of CH₂Cl₂. Crystal structure data for **2**: Formula C₃₆H₂₈O₂, M= 492.61, crystal dimensions 0.15 x 0.15 x 0.15 mm, monoclinic, space group P21/c, $a = 11.3010$ (23) Å, $b = 11.8081$ (24) Å, $c = 19.1846$ (40) Å, $\beta = 96.897$ (3), $V = 2541.53$ Å³, $Z = 4$, $\rho_{\text{calcd}} = 1.2872$ g cm⁻³. (see Figure 19)

	Phenyl-phenyl [Å] (centroids to centroids)	3.498
25	Splaying angle between peri-aryl planes [°]	19.48
	Torsion angle [°]	0.28

5. Enantioselective recognition

We now have found that **2** can be used for enantioselective recognition of a wide range of amines. The CD spectra of the enantiomers of **2** are shown in Figure 15. Upon addition of a sufficiently strong base, the corresponding bisphenolate having a remarkably different CD spectrum is formed. Addition of enantiomers of chiral amines then regenerates the original CD spectrum. This proceeds with high stereoselectivity and therefore allows enantioselective CD sensing of the amine used. Proof-of-concept data were obtained with different analytes, including 1-phenylethylamine, Figure 20 (shows titration curve with (-)-**2**); Figure 21 (shows titration curve with (+)-**2**). (For CD analysis of amines with chiral metal complexes, see: (a) Nieto, S.; Lynch, V. M.; Anslyn, E. V.;

- 5 Kim, H.; Chin, J. *Org. Lett.* 2008, 10, 5167. (b) Nieto, S.; Lynch, V. M.; Anslyn, E. V.; Kim, H.; Chin, J. *J. Am. Chem. Soc.* 2008, 130, 9232. (c) Nieto, S.; Dragna, J. M.; Anslyn, E. V. *Chem. Eur. J.* 2010, 16, 227.)

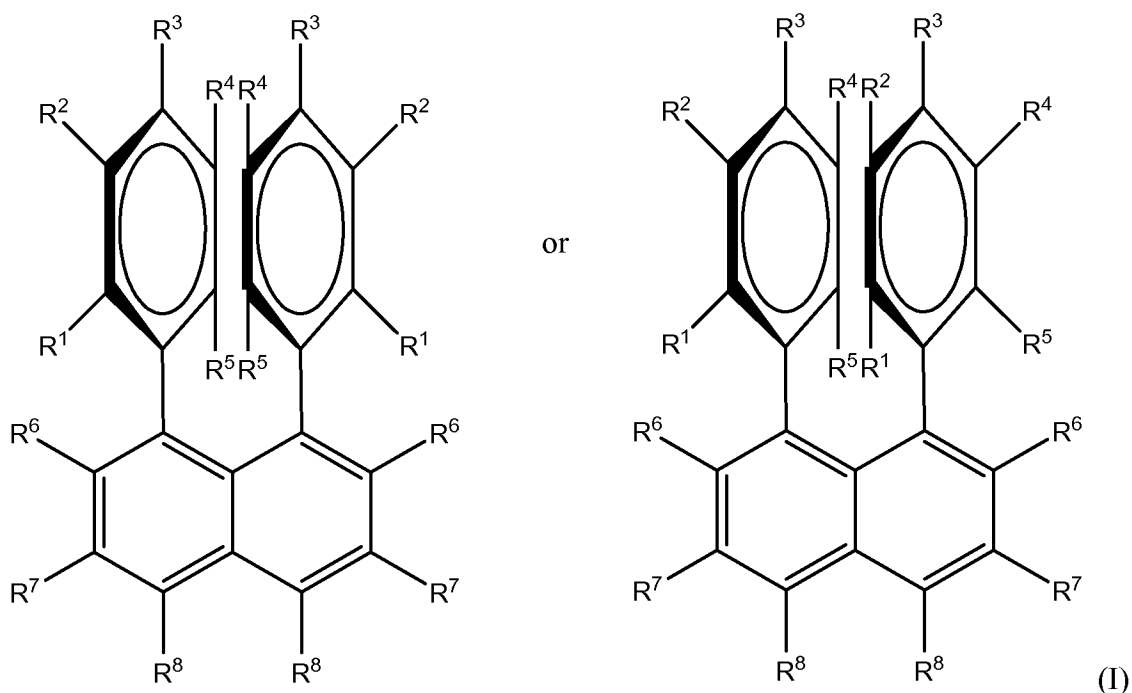
10 The following conditions have been optimized in terms of reaction time, solvent, concentration and equivalents. Prior to each use, the CD instrument was purged with nitrogen for 20 minutes. Spectra were collected between 240 and 600 nm with a standard sensitivity of 100 mdeg, a data pitch of 0.5 nm, a band width of 0.5 nm, a scanning speed of 1000 nm s⁻¹ and a response of 0.5 s using a quartz cuvette (1 cm path length). The concentration of (-)-**2** was 9.45 x 10⁻⁵ M in ACN. To 2000 µL of (-)-**2**, 4 equivalents of
15 Na^tBuO (0.521 M in DMSO) were added to generate the bisphenoxide of (-)-**2**, which was titrated with several amine substrates until the original bisphenol CD spectrum was recovered. Ellipticities at 290 nm are then plotted against the concentration of added amine enantiomers. Data obtained with (*R*)-enantiomers are represented by blue triangles and those obtained with (*S*)-enantiomer are shown in red (see Figures 22-30)

20 In conclusion, chiral bisphenol **2** has been prepared and used for enantioselective sensing of a wide range of amines. This study shows the potential of this class of compounds for chiral recognition and it points to the use in asymmetric catalysis.

25 Having thus described in detail various embodiments of the present invention, it is to be understood that the invention defined by the above paragraphs is not to be limited to particular details set forth in the above description as many apparent variations thereof are possible without departing from the spirit or scope of the present invention.

CLAIMS

1. An atropisomeric 1,8-bisphenolnaphthalene compound of formula (I):



wherein

- (a) R^1 and R^5 are independently selected from the group consisting of hydrogen, alkyl, alkenyl, alkynyl, cycloalkyl, aryl, benzyl, heterocyclyl, heteroaryl, chloro, bromo, iodo, acyl, amino, amido, azido, cyano, formyl, carbamoyl, $-SF_5$, nitro, $-OR'$, $-NR''R'''$, $-SR''''$, $-P(O)(OR^A)$, $-P(OR^A)_2$, and $P(R^B)_2$, wherein

R' is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical, $-P(O)(OR^A)$, $-P(OR^A)_2$, and $P(R^B)_2$; or OR' is an optionally substituted saturated or unsaturated, nonaromatic or aromatic heterocyclic radical;

R'' and R''' are each independently of one another H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical; or $NR''R'''$ is an optionally substituted saturated or unsaturated, nonaromatic or aromatic heterocyclic radical;

R^{''''} are each independently of one another H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical; or SR^{''} is an optionally saturated or unsaturated, nonaromatic or aromatic heterocyclic radical;

R^A is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

R^B is H, halogen, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

wherein at least one of R¹ and R⁵ on each phenyl ring is not hydrogen;

(b) R², R³, R⁴, R⁶, R⁷ and R⁸ are independently selected from the group consisting of hydrogen, alkyl, alkenyl, alkynyl, cycloalkyl, aryl, benzyl, heteroalkyl, heterocyclyl, heterocycloalkyl, heteroaryl, halogen, acyl, amino, amido, azido, alkyliminyl (-C=NH-alkyl), carboxy, cyano, formyl, carbamoyl, -SF₅, nitro, OR['], NR^{''}R^{'''}, SR^{'''}, -P(O)(OR^A), -P(OR^A)₂, and P(R^B)₂, wherein

R['] is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical, -P(O)(OR^A), -P(OR^A)₂, and P(R^B)₂; or OR['] is an optionally substituted saturated or unsaturated, nonaromatic or aromatic heterocyclic radical;

R^{''} and R^{'''} are each independently of one another H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarbonoxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical; or NR^{''}R^{'''} is an optionally

substituted saturated or unsaturated, nonaromatic or aromatic heterocyclic radical;

R'''' are each independently of one another H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarboxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical; or SR'' is an optionally saturated or unsaturated, nonaromatic or aromatic heterocyclic radical;

R^A is H, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarboxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

R^B is H, halogen, alkyl, alkenyl, alkynyl, cycloalkyl, cycloalkenyl, an (acyclic or cyclic) hydrocarbon radical, hydrocarboxy radical, a heterocyclyl radical, heterocyclyloxy radical, heterocyclylthio radical or heterocyclylamino radical;

wherein at least one of R² and R⁴ on each phenyl ring is not hydrogen.

2. An atropisomeric 1,8-bisphenolnaphthalene of claim 1, wherein

(a) R¹ and R⁵ are independently selected from the group consisting of hydrogen, (C₁-C₄) alkyl, phenyl, chloro, bromo and iodo, wherein at least one of R¹ and R⁵ on each phenyl ring is not hydrogen;

(b) R² and R⁴ are independently selected from the group consisting of hydrogen, (C₁-C₄) alkyl, (C₂-C₄) alkenyl, (C₂-C₄) alkynyl, (C₃-C₆) cycloalkyl, phenyl, (C₁-C₆) acyl, and formyl, wherein at least one of R² and R⁴ on each phenyl ring is not hydrogen.;

(c) R³ is independently PR'₂, NR'₂ or OR' wherein each R' is independently H or (C₁-C₄) alkyl, and

(d) R^6 , R^7 and R^8 are independently selected from the group consisting of hydrogen, (C₁-C₄) alkyl, hydroxyl, (C₁-C₄) alkoxy, halogen, NH₂, amido, azido, cyano, formyl, carbamoyl, nitro.

3. An atropisomeric 1,8-bisphenolnaphthalene of claim 2, wherein

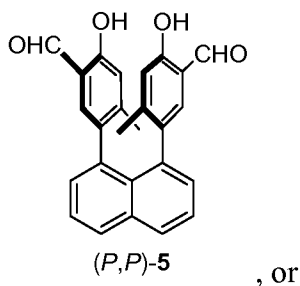
(a) R^1 and R^5 are independently H and methyl, wherein at least one of R^1 and R^5 on each phenyl ring is not hydrogen;

(b) R^2 and R^4 are independently phenyl or formyl, wherein at least one of R^2 and R^4 on each phenyl ring is not hydrogen;

(c) R^3 is independently OR' wherein R' is H or methyl; and

(d) R^6 , R^7 and R^8 are independently selected from the group consisting of hydrogen and methyl.

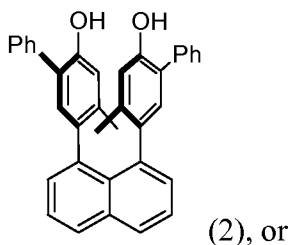
4. An atropisomeric 1,8-bisphenolnaphthalene compound of claim 3, wherein compound has the structure of formula 5 below:



, or

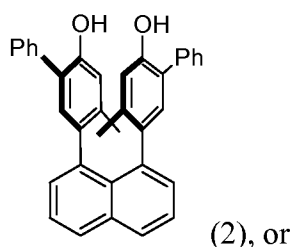
its enantiomer.

5. An atropisomeric 1,8-bisphenolnaphthalene compound of claim 3, wherein compound has the structure of formula 2 below:



its enantiomer.

6. An atropisomeric 1,8-bisphenolnaphthalene compound of claim 1, 2, 3, 4 or 5, where in the compound is stable to racemization at room temperature (20-25°C).
7. A method of providing enantiomeric recognition between stereoisomers of a chiral compound, which comprises of adding an atropisomeric 1,8-bisphenolnaphthalene compound of claim 1, 2, 3, 4 or 5 to a solution containing a racemic or diastereomeric mixture of the chiral compound, wherein the chiral compound is an amine, amino alcohol amino acid or alcohol.
8. The method of claim 7, wherein the atropisomeric 1,8-bisphenolnaphthalene compound has the structure of formula 2 below:



its enantiomer.

9. The method of claim 7, wherein the chiral amine is selected from the group consisting of 1-phenylethylamine, 3,3-dimethylbutan-2-amine, 3-dimethylbutan-2-amine, cyclohexane-1,2-diamine, 1,2-diphenylethane-1,2-diamine, 1,1-diphenylpropan-2-amine, 1-cyclohexylethanamine, 2,6,6-trimethylbicyclo[3.1.1]heptan-3-amine, and heptan-2-amine.

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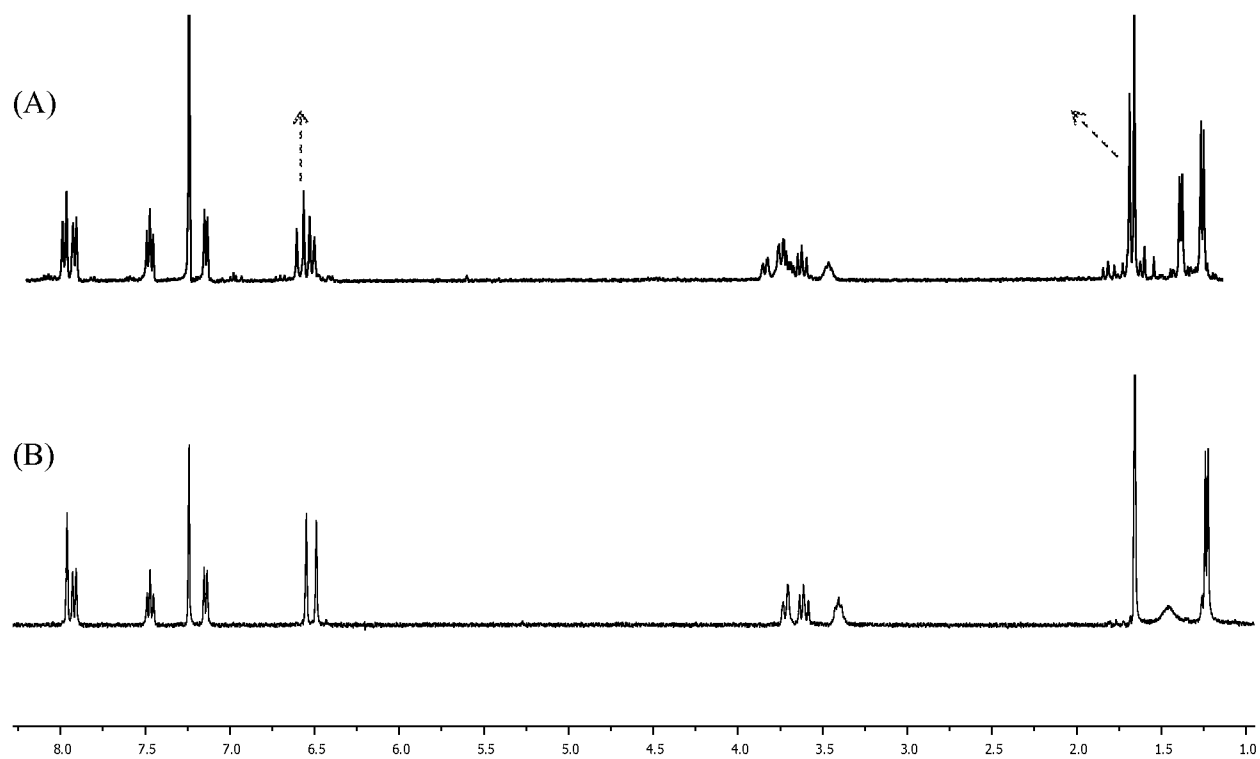


Figure 1

2 of 30

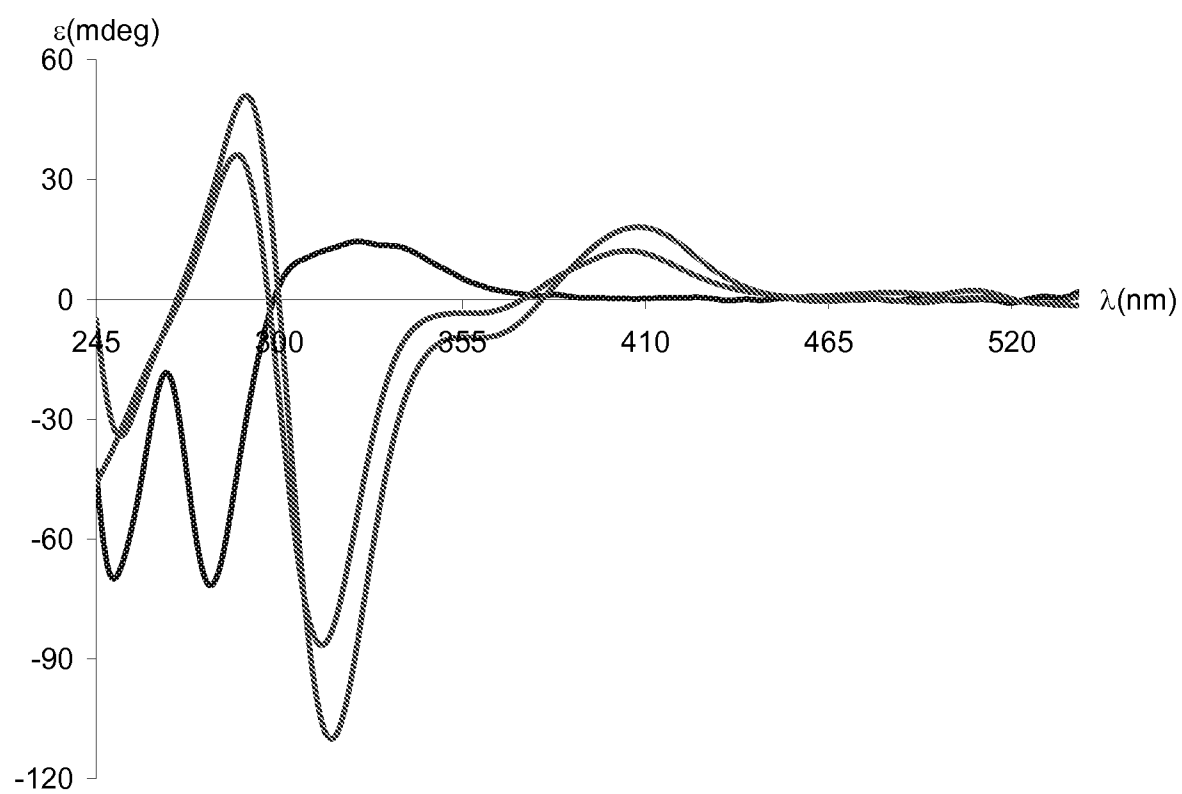


Figure 2

3 of 30

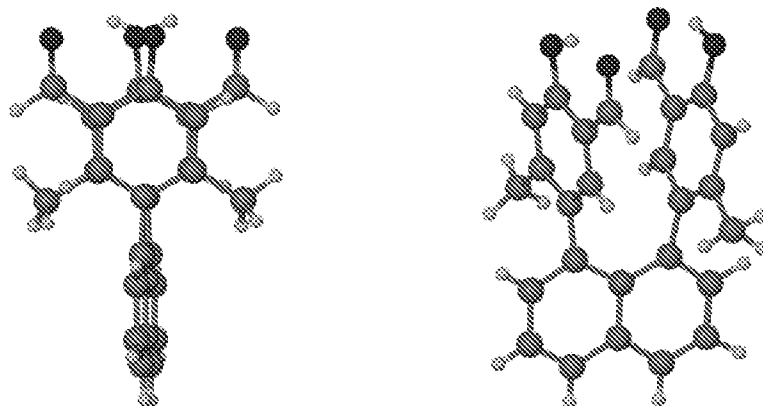


Figure 3

4 of 30

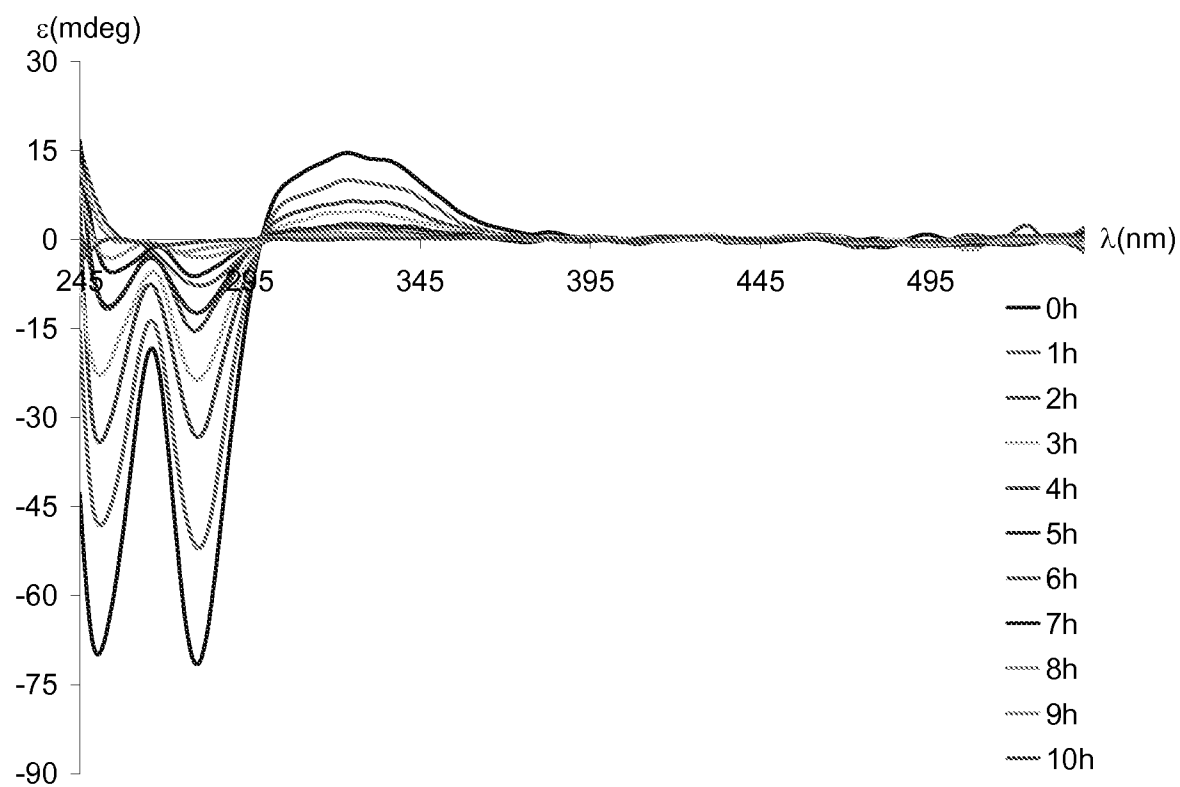


Figure 4

5 of 30

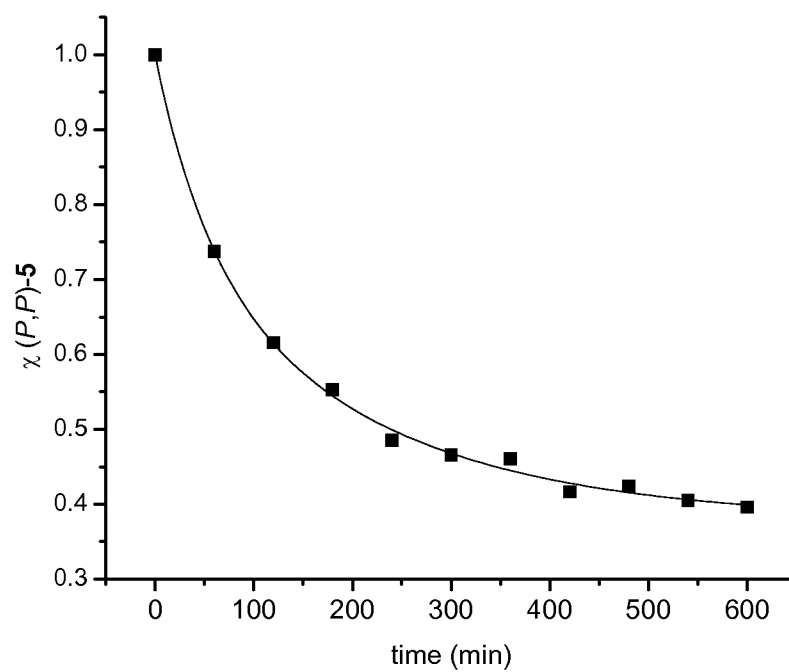


Figure 5

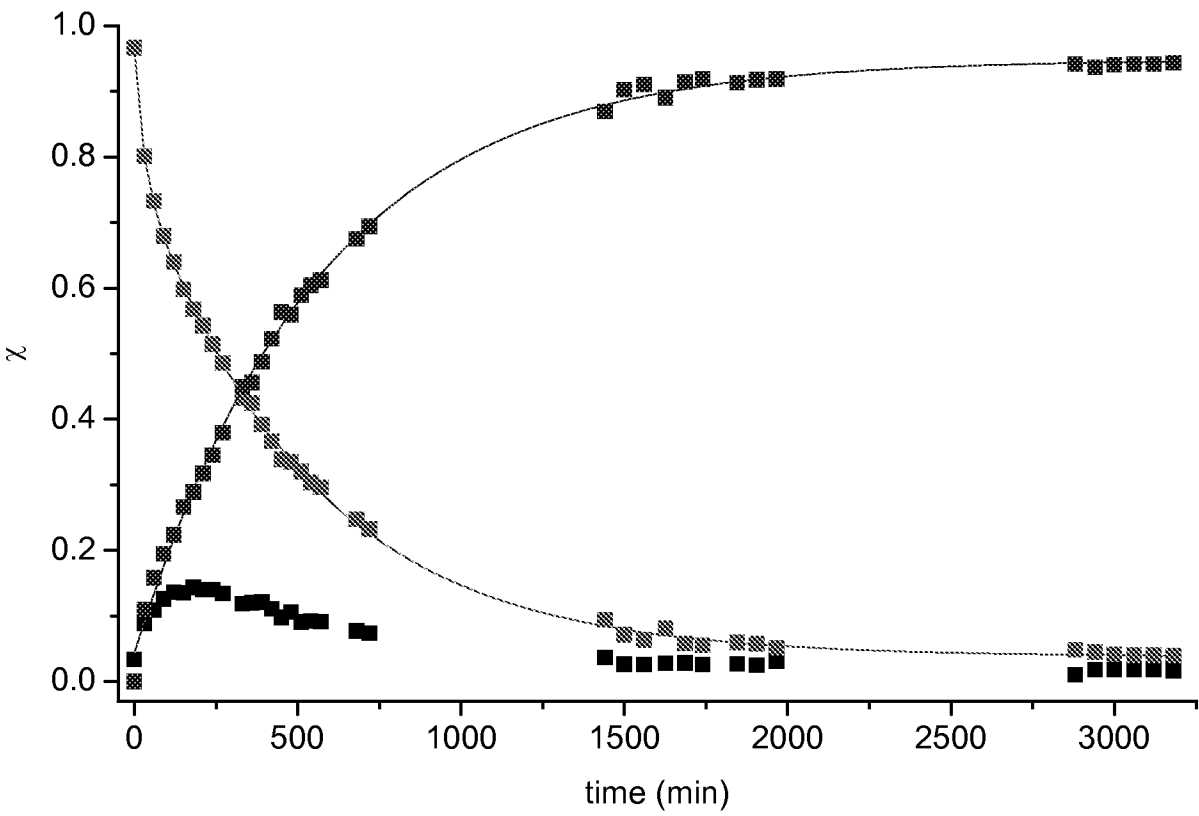


Figure 6

7 of 30

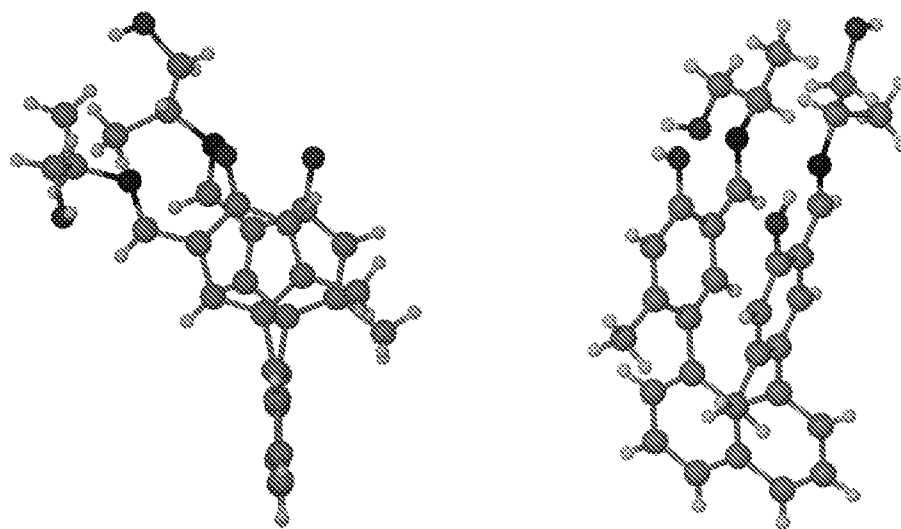
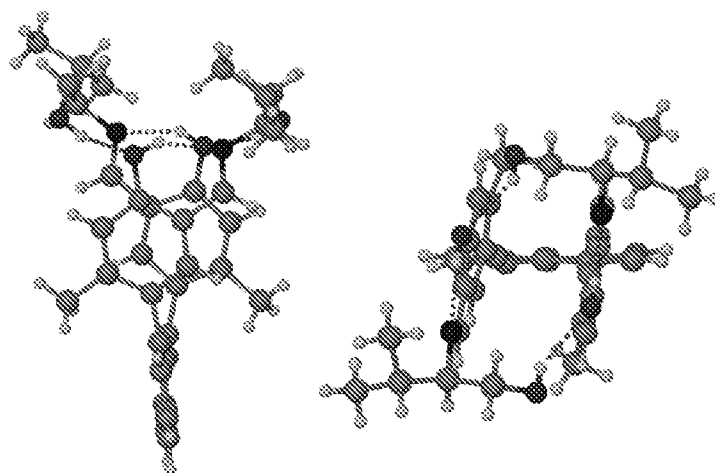


Figure 7

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**Figure 8**

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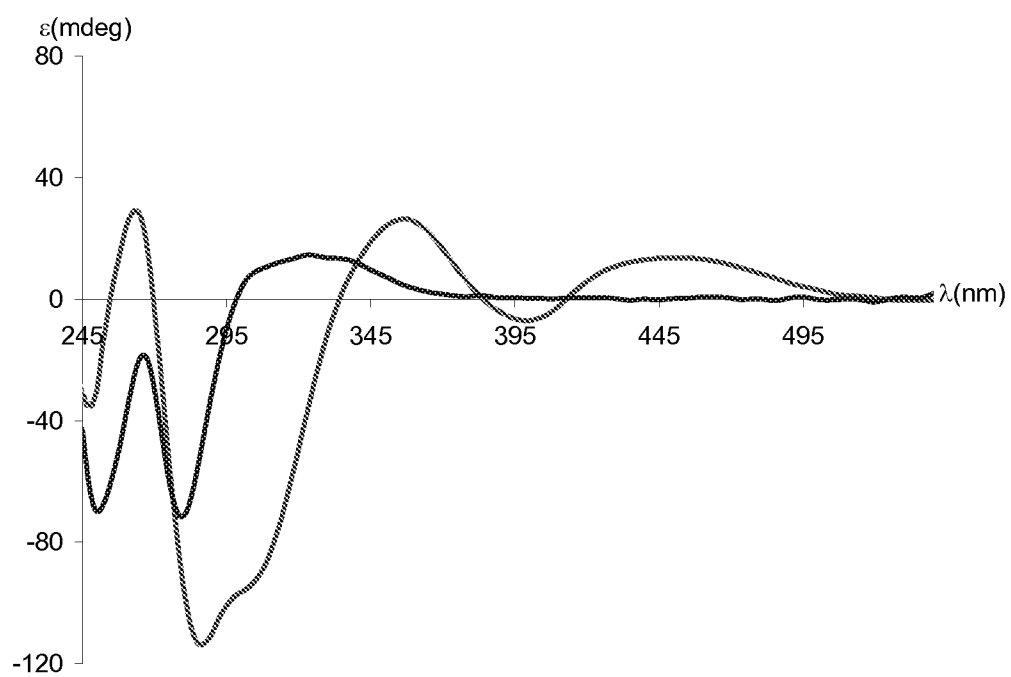


Figure 9

10 of 30

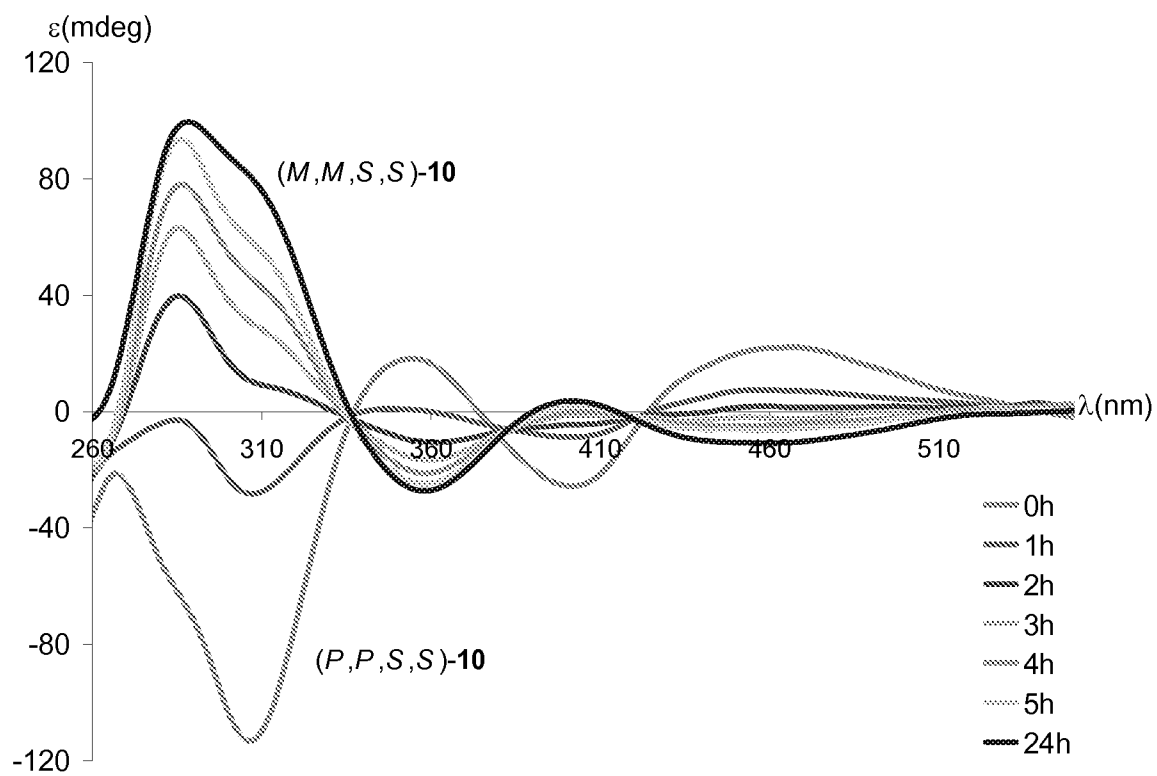


Figure 10

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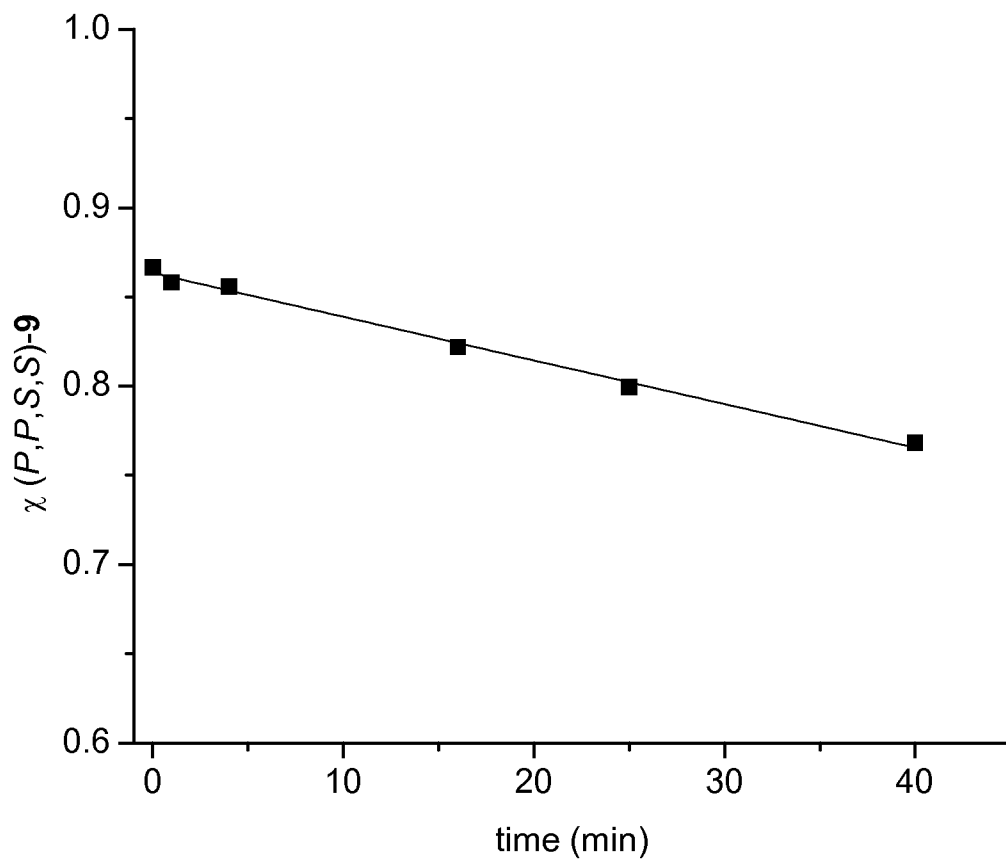


Figure 11

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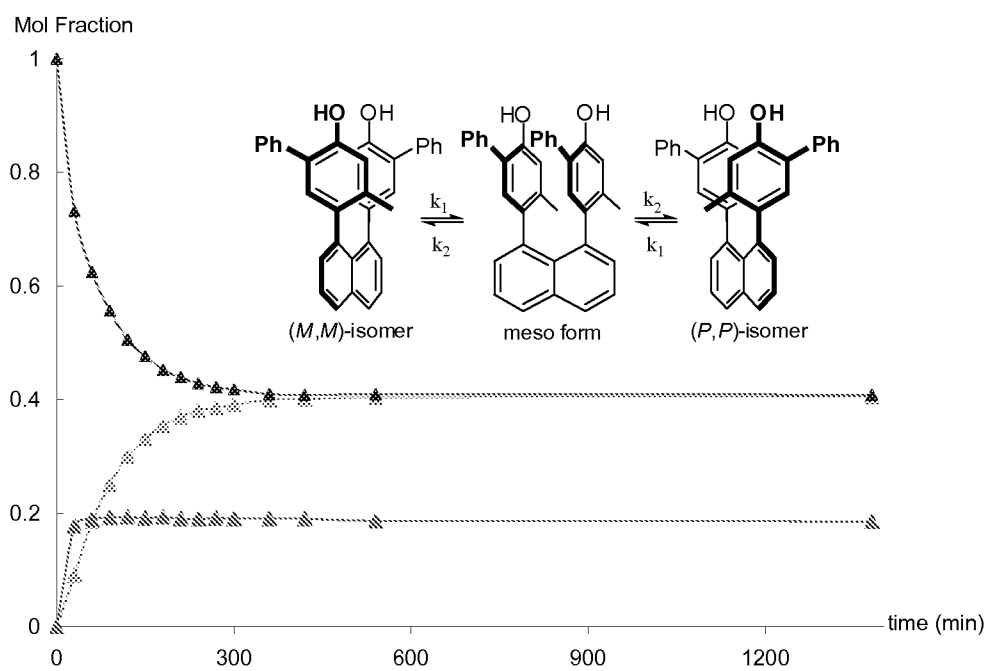


Figure 12

13 of 30

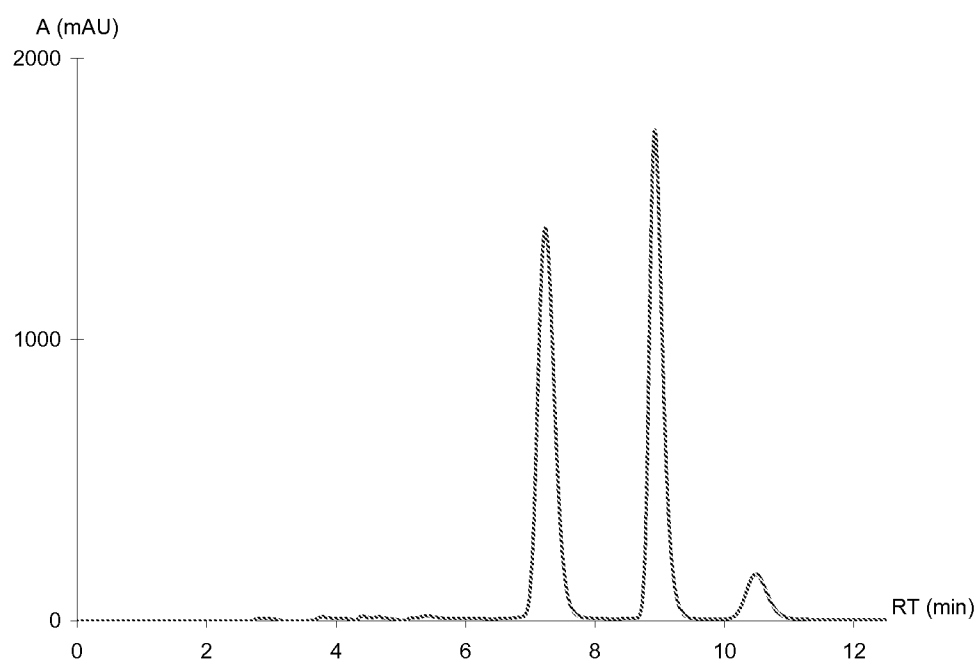


Figure 13

14 of 30

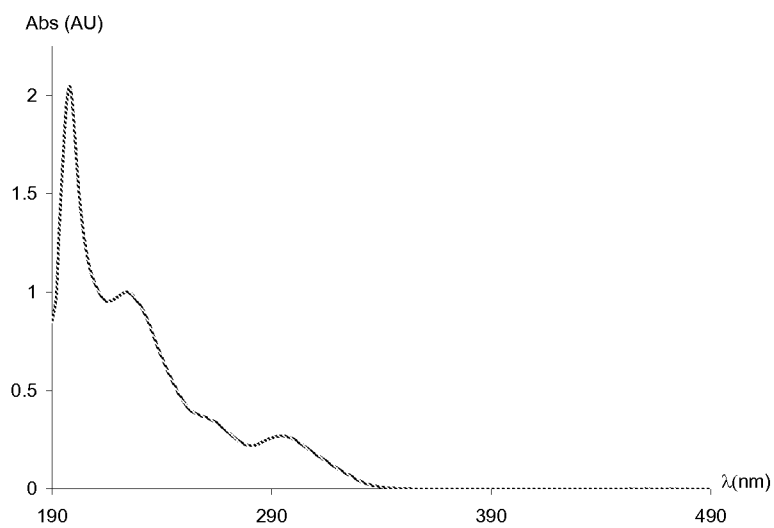


Figure 14a

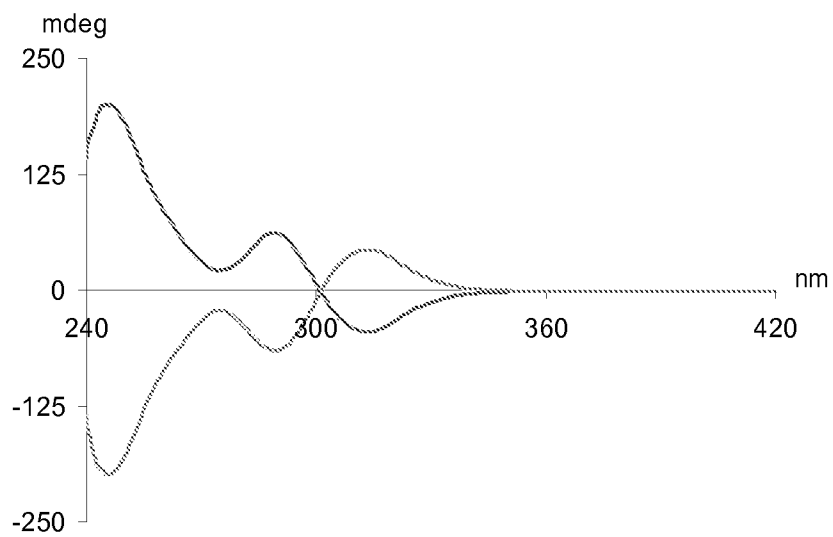


Figure 14b

15 of 30

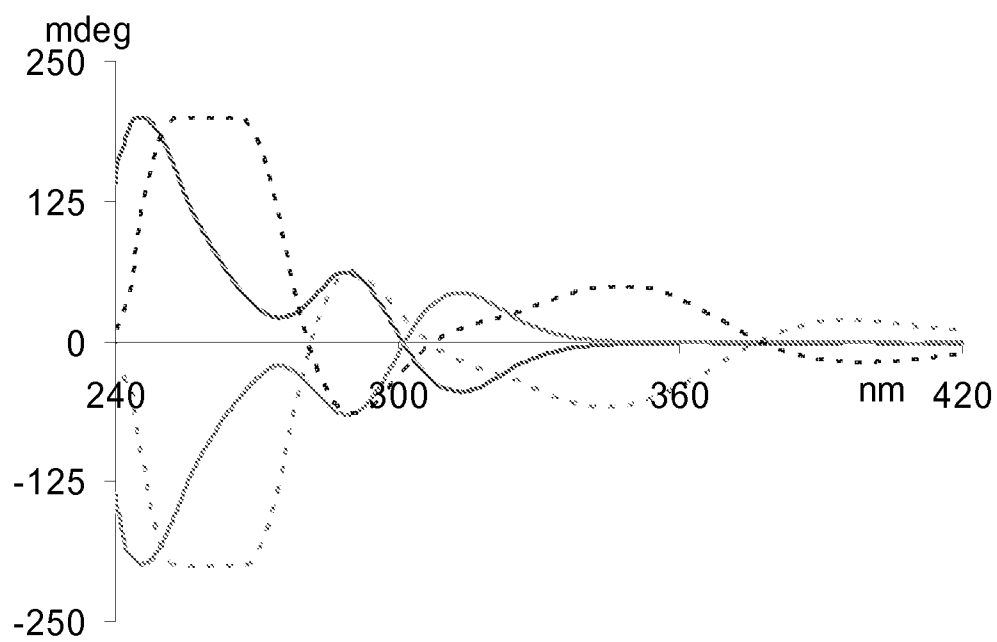


Figure 15

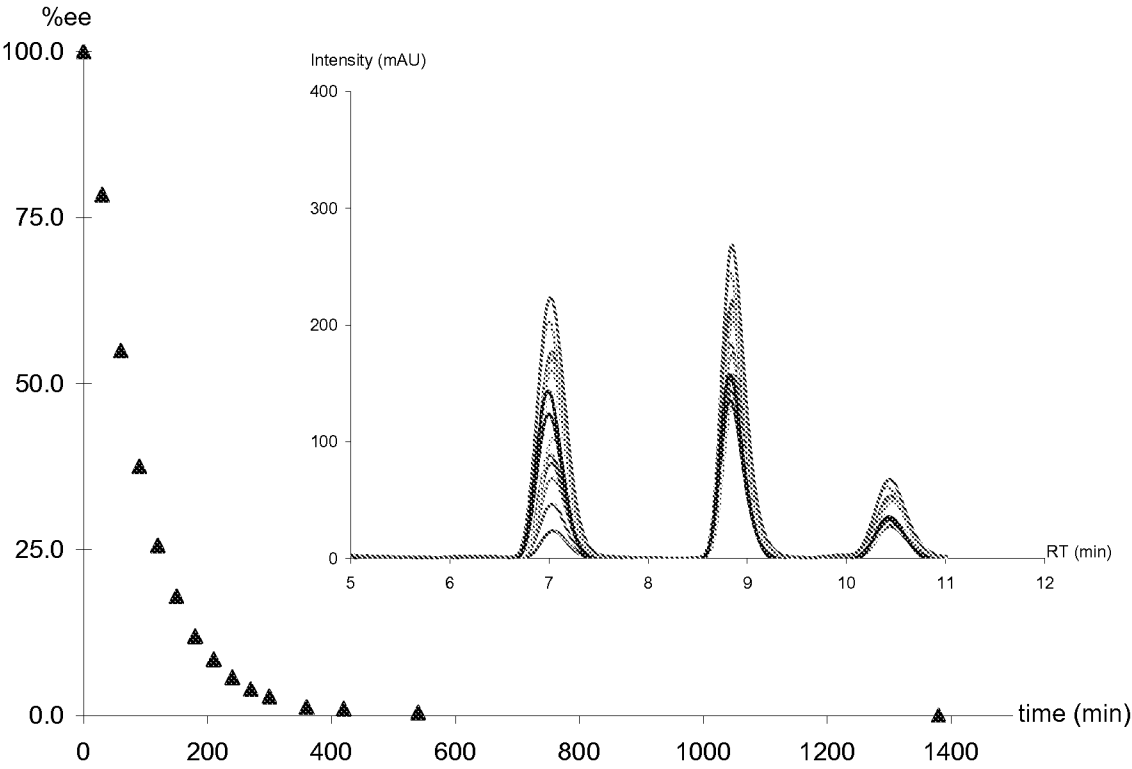


Figure 16

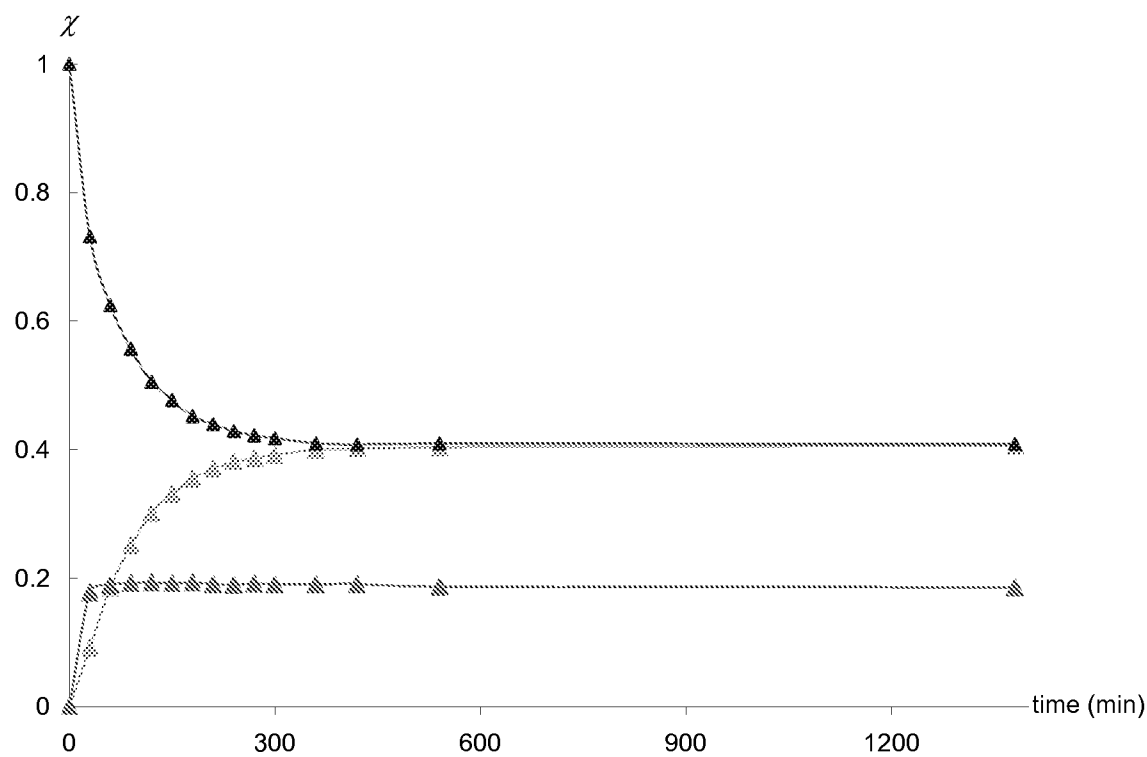


Figure 17

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Figure 18

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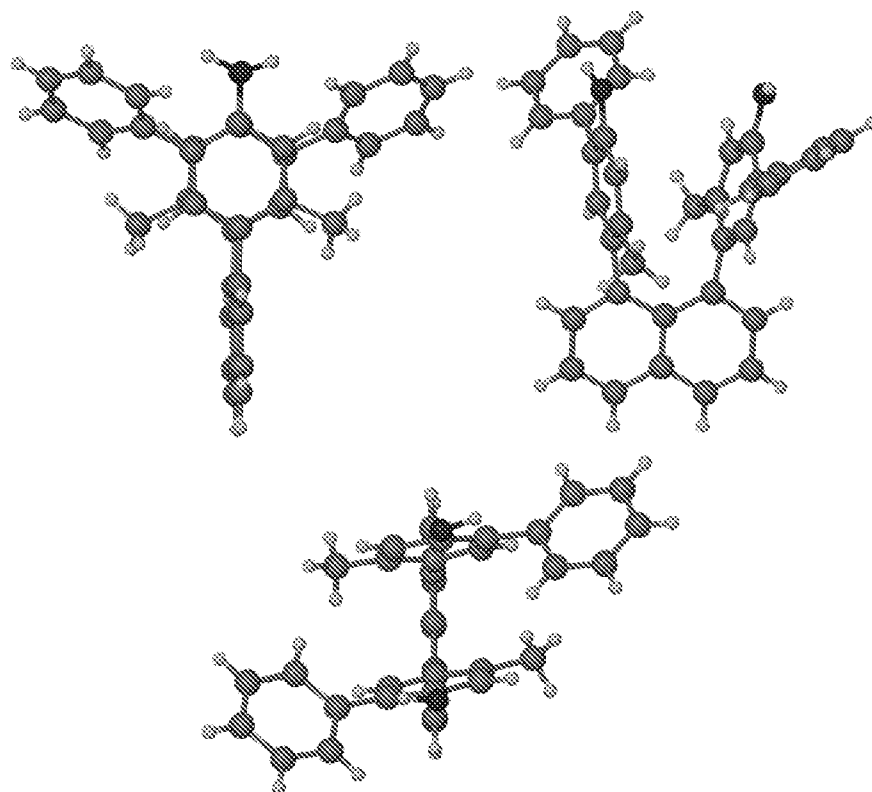


Figure 19

20 of 30

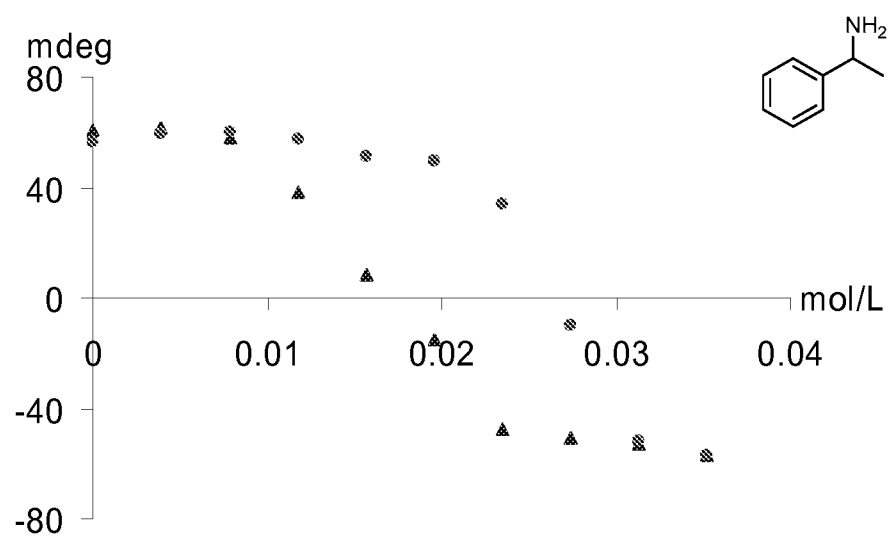


Figure 20

21 of 30

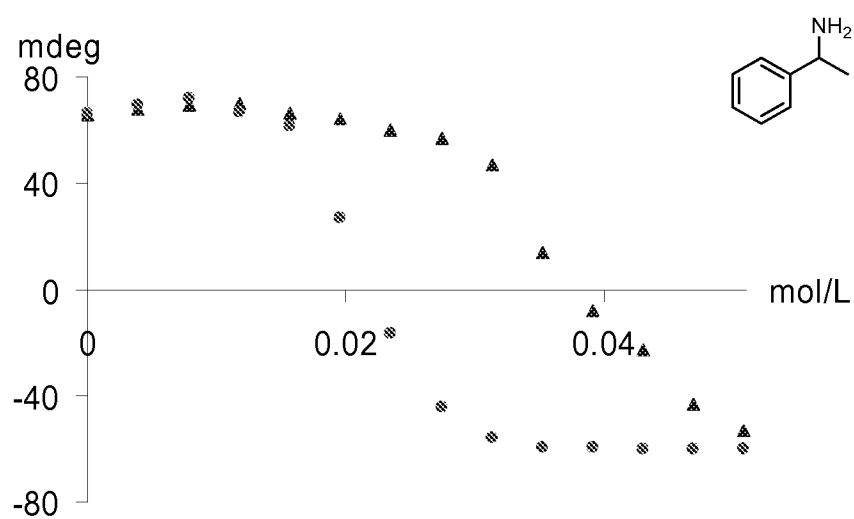


Figure 21

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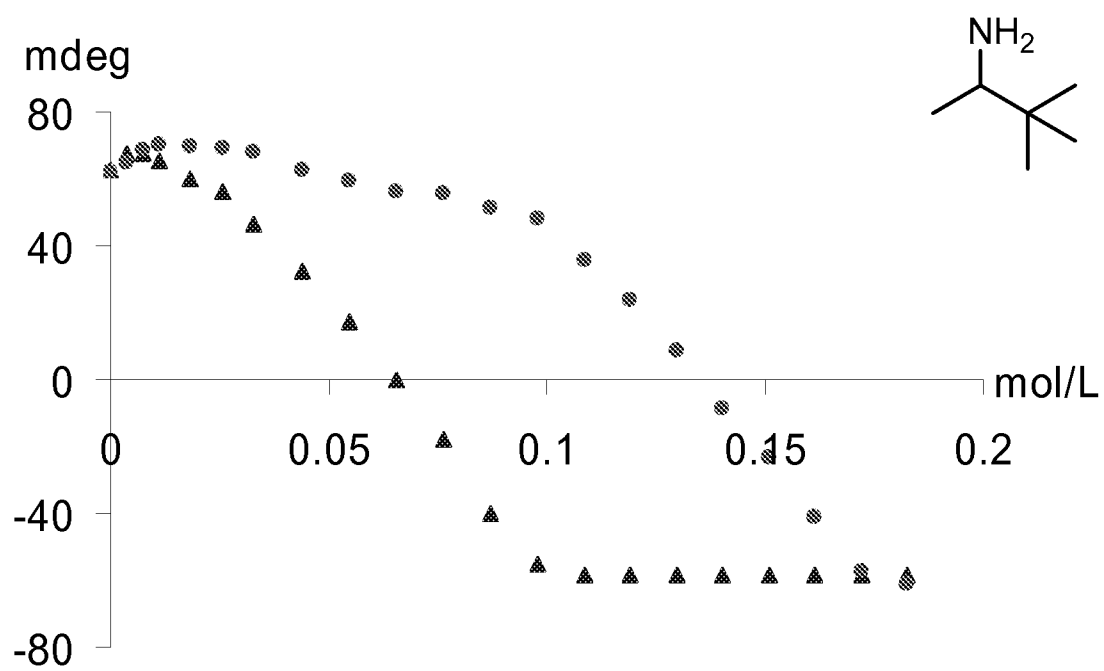


Figure 22

23 of 30

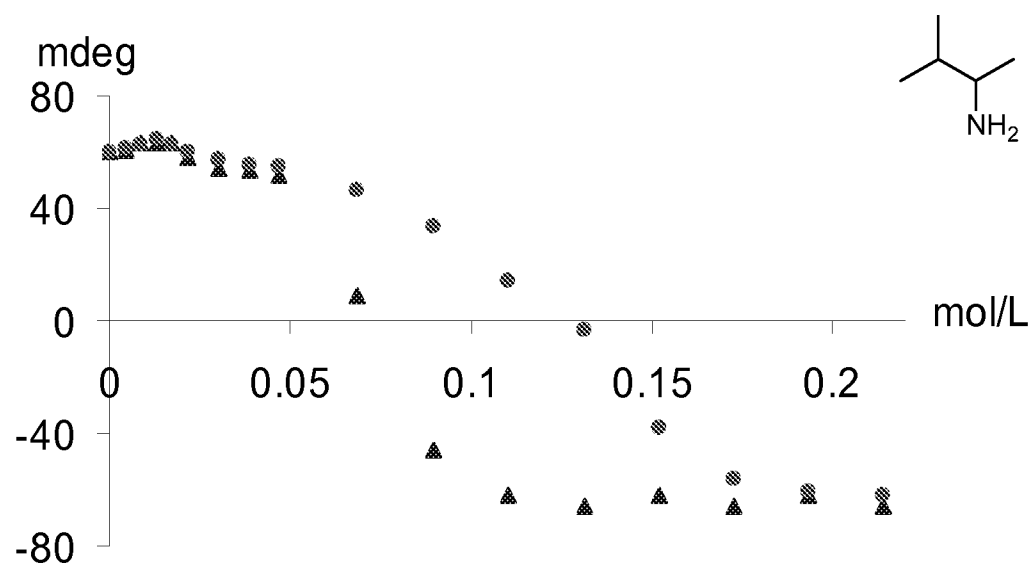


Figure 23

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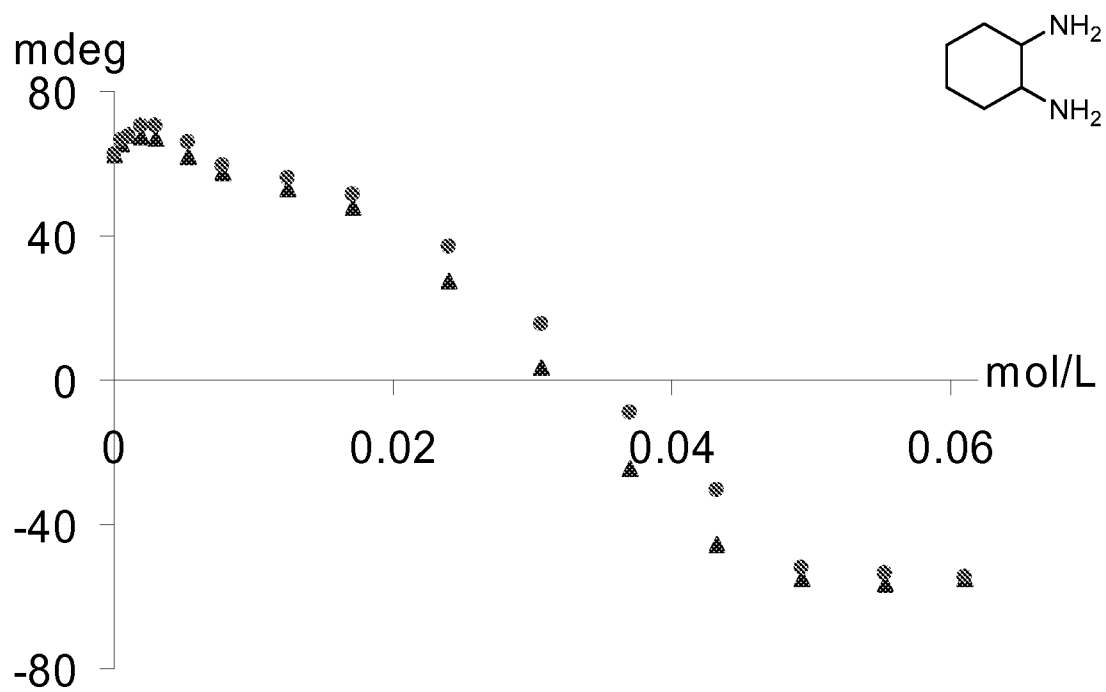


Figure 24

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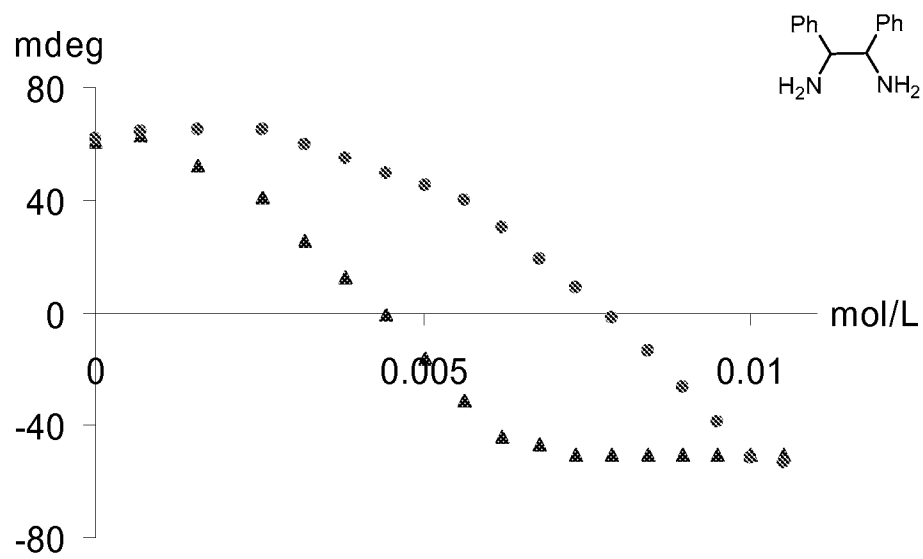


Figure 25

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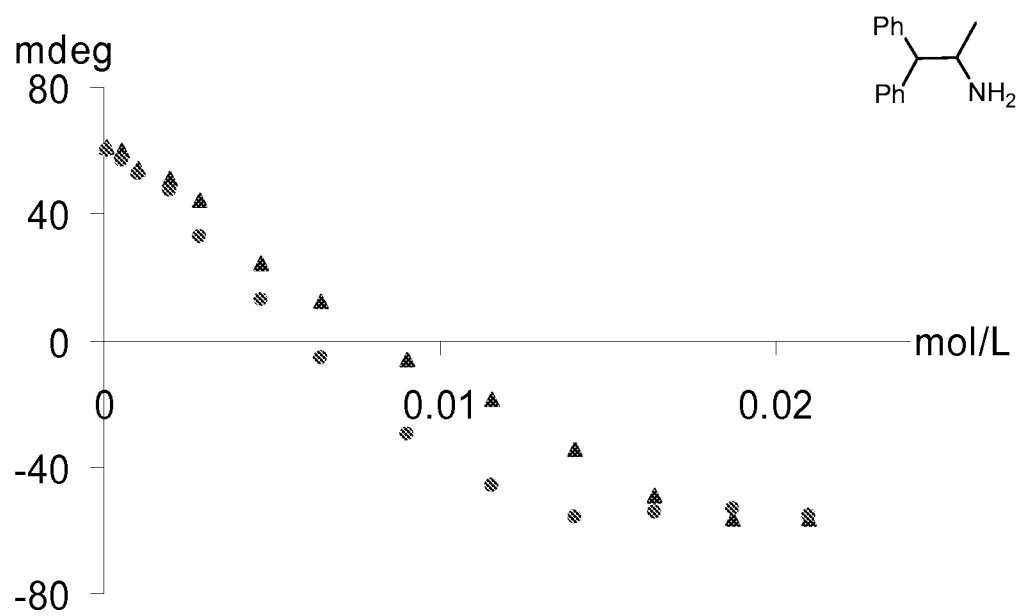


Figure 26

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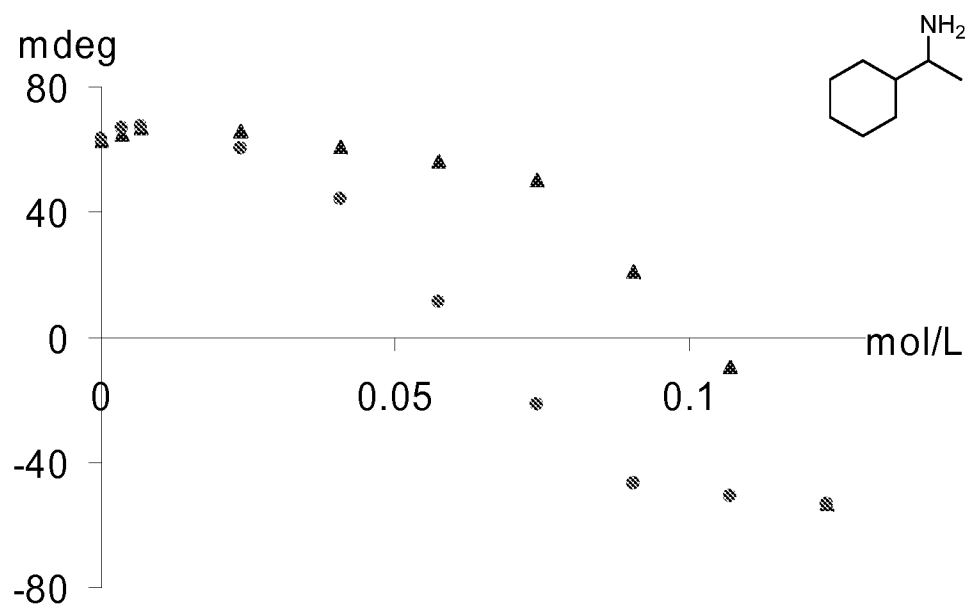


Figure 27

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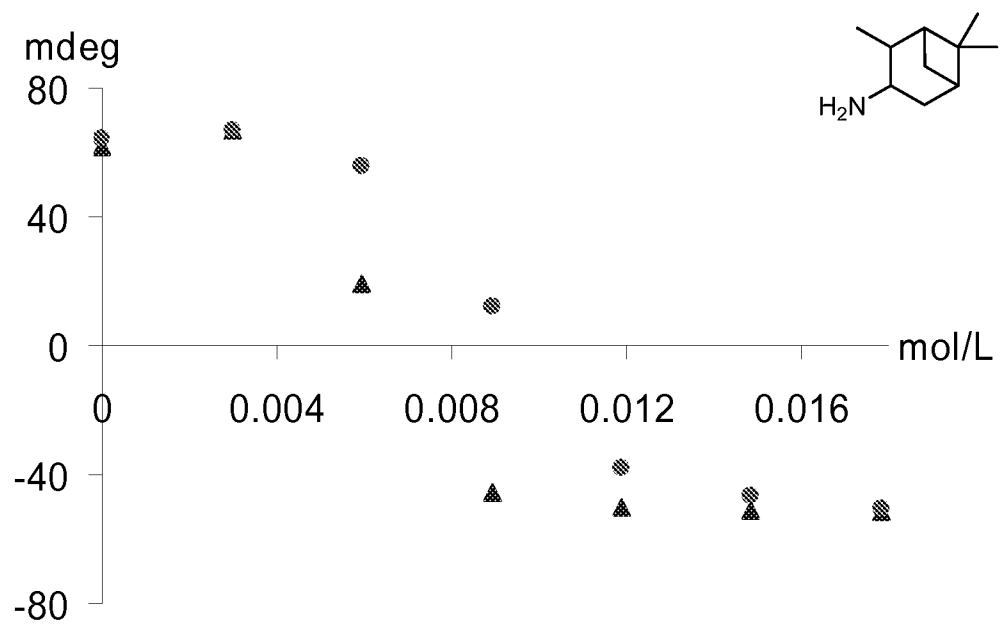


Figure 28

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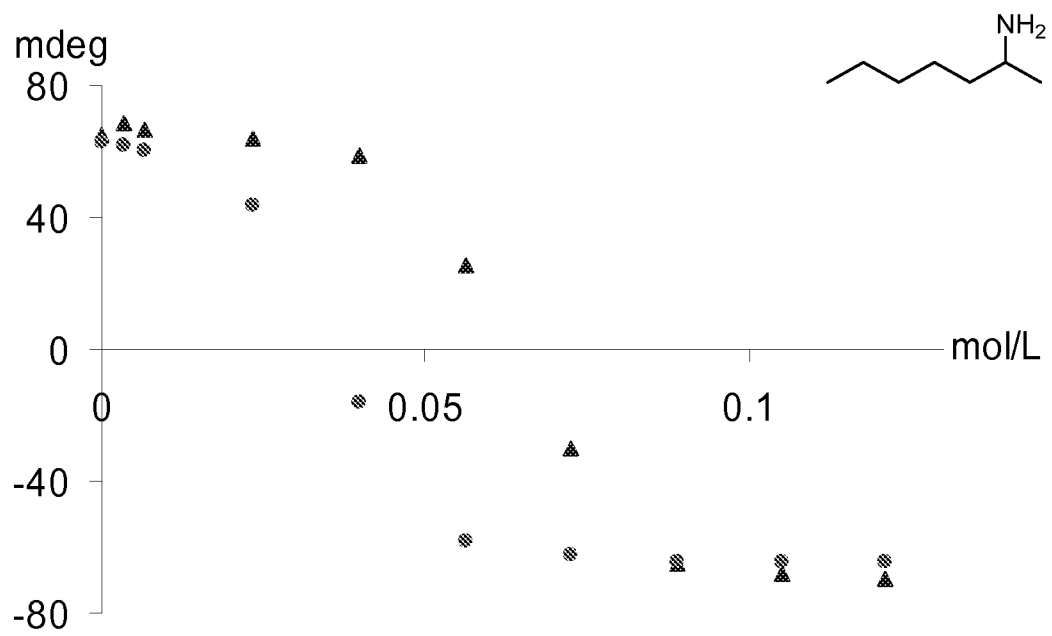


Figure 29

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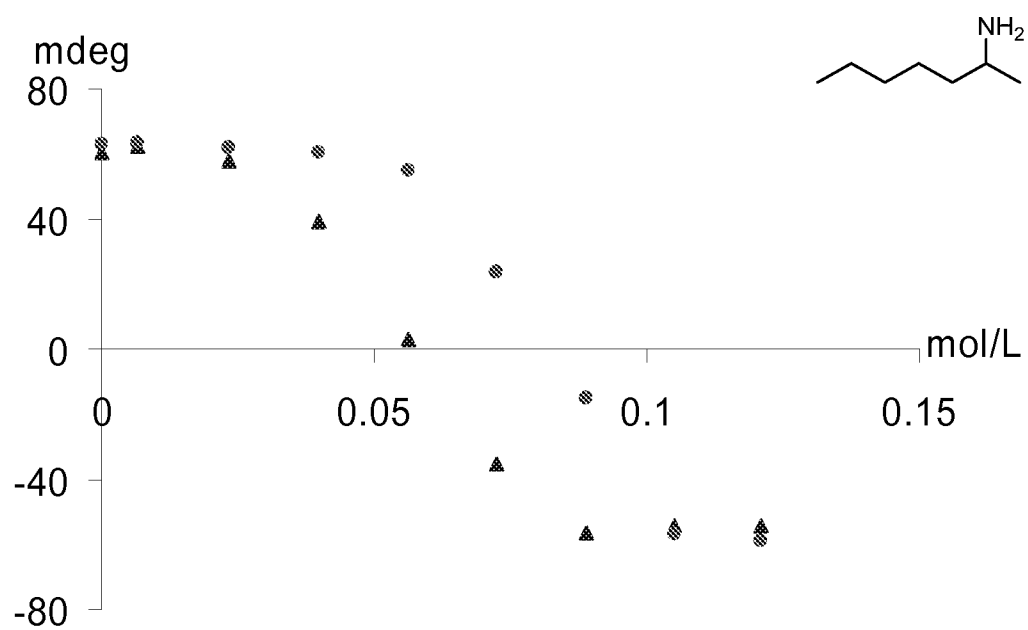


Figure 30