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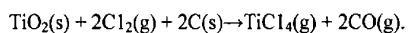
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(54) Title: THERMAL AND ELECTROCHEMICAL PROCESS FOR METAL PRODUCTION

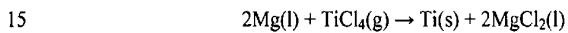
(57) Abstract: A system for purification of high value metals comprises an electrolytic cell in which an anode formed of a composite of a metal oxide of the metal of interest with carbon is electrochemically reduced in a molten salt electrolyte.

## THERMAL AND ELECTROCHEMICAL PROCESS FOR METAL PRODUCTION

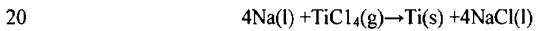
The properties of titanium have long been recognized as a light, strong, and 5 corrosion resistant metal, which has lead to many different approaches over the past few decades to extract titanium from its ore. These methods were summarized by Henrie [1]. Despite the many methods investigated to produce titanium, the only methods currently utilized commercially are the Kroll and Hunter processes [2, 3]. These processes utilize titanium tetrachloride ( $TiCl_4$ ) which is produced from the 10 carbo-chlorination of a refined titanium dioxide ( $TiO_2$ ) according to the reaction:



In the Kroll process [2]  $TiCl_4$  is reduced with molten magnesium at  $\approx 800^\circ C$  in an atmosphere of argon. This produces metallic titanium as a spongy mass according to the reaction:



from which the excess Mg and  $MgCl_2$  is removed by volatilization, under vacuum at  $\approx 1000^\circ C$ . The  $MgCl_2$  is then separated and recycled electrolytically to produce Mg as the reductant to further reduce the  $TiCl_4$ . In the Hunter process [3,4] sodium is used as a reductant according to the reaction:



The titanium produced by either the Kroll or Hunter processes must not only be separated from the reductant halide by vacuum distillation and/or leaching in acidified solution to free the titanium sponge for further processing to useful titanium forms, but also require the recycling of the reductant by electrolysis. Because of 25 these multiple steps the resultant titanium is quite expensive which limits its use to cost insensitive applications.

The US Bureau of Mines performed extensive additional investigations [1,5-8] to improve the Kroll and Hunter processes. Many other processes have been investigated that include plasma techniques [9-13], molten chloride salt electrolytic 30 processes [14], molten fluoride methods [15], the Goldschmidt approach [16], and alkali metal-calcium techniques [17]. Other processes investigated without measurable success have included aluminum, magnesium, carbothermic and carbo-nitrothermic reduction of  $TiO_2$  and plasma reduction of  $TiCl_4$  [18]. Direct reduction

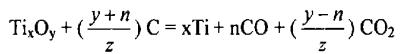
of  $\text{TiO}_2$  or  $\text{TiCl}_4$  using mechanochemical processing of ball milling with appropriate reductants of Mg or calcium hydride ( $\text{CaH}_2$ ) also have been investigated [19] without measurable success. Kroll, who is considered as the father of the titanium industry [20] predicted that titanium will be made competitively by fusion electrolysis but to

5 date, this has not been realized.

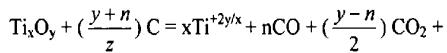
An electrolytic process has been reported [21] that utilizes  $\text{TiO}_2$  as a cathode and carbon or graphite as the anode in a calcium chloride electrolyte operated at 900°C. By this process, calcium is deposited on the  $\text{TiO}_2$  cathode, which reduces the  $\text{TiO}_2$  to titanium and calcium oxide. However, this process is limited by diffusion of 10 calcium into the  $\text{TiO}_2$  cathode and the build-up of calcium oxide in the cell, which limits operating time to remove the calcium oxide or replacement of the electrolyte. Also the  $\text{TiO}_2$  cathode is not fully reduced which leaves contamination of  $\text{TiO}_2$  or reduced oxides such as  $\text{TiO}$ , mixed oxides such as calcium titanate as well as 15 titanium carbide being formed on the surface of the cathode thus also contaminating the titanium. Thus, current  $\text{TiO}_2$  cathode electrolytic processes are no more commercially viable than earlier electrolytic processes.

The instant invention is a combination of a thermal and an electrochemical process, which utilizes a carbon or composite anode containing a metal oxide of a metal of interest, as a feed electrode. As used herein the term "carbon" is meant to 20 include carbon in any of its several crystalline forms including, for example, graphite. For example, for producing purified titanium, the feed should comprise  $\text{TiO}_2$  which may be high purity, rutile, synthetic rutile, ilmenite or other source of titanium, mixed with a source of carbon and pressed together with or without a binder that also may be a source of carbon on pyrolysis to form a  $\text{TiO}_2$ -C composite 25 green electrode or billet. The  $\text{TiO}_2$ -C composite billet is then heated, in the absence of air to avoid oxidation of the carbon component, to a temperature sufficient to reduce the plus four valence of the titanium in the  $\text{TiO}_2$  to a lower valence. The temperature of heating and time at temperature will determine the reduced oxide stoichiometry of the titanium oxide which may be expressed as  $\text{Ti}_x\text{O}_y$  where the ratio 30 of  $y/x$  can be 0 to equal or less than 2 and  $y$  balances the valence charge of the titanium species. Some examples of reduced titanium oxide compounds include  $\text{TiO}$ ,  $\text{Ti}_2\text{O}_3$ ,  $\text{Ti}_3\text{O}_5$ , and  $\text{Ti}_4\text{O}_7$  and mixtures thereof. Sufficient residual carbon needs to remain after the thermal reduction step or can be added separately to

stoichiometrically react with the reduced titanium oxide to electrochemically produce titanium at the cathode and CO<sub>2</sub> and/or CO at the anode. The reduced titanium state oxide composite anode overall general reactions are:

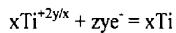


5 at the anode:



zye<sup>-</sup>

where 2y/x is the oxide state of the titanium in the electrolyte,  
10 at the cathode:



#### Summary of the Invention

In an embodiment of the invention there is provided a method for the production of titanium metal which comprises electrochemically dissolving, in a 15 molten salt electrolyte, an anode formed of a titanium suboxide/carbon composite, and reducing the dissolved titanium suboxide, at a cathode, to titanium metal.

In an embodiment of the invention there is provided a method for the production of purified titanium from rutile ore which comprises electrowinning from an anode formed of a mixture of titanium suboxide/carbon composite in a molten salt 20 electrolyte, and depositing purified titanium onto a cathode.

In an embodiment of the invention there is provided a method for the production of purified titanium which comprises electrochemically dissolving an anode formed of a titanium suboxide/carbon composite in a molten salt electrolyte, and electrochemically reducing the dissolved titanium suboxide to purified titanium 25 metal.

In an embodiment of the invention there is provided a method for the direct production of titanium metal in a particulate state which comprises electrochemically dissolving an anode, formed of a titanium suboxide/carbon composite, in a molten salt electrolyte in an electrochemical cell, and electrochemically reducing the 30 dissolved titanium suboxide to particulate titanium metal.

In an embodiment of the invention there is provided a method for the production of titanium metal which comprises electrochemically dissolving, in a molten salt electrolyte, an anode formed of a titanium oxide/carbon composite, wherein the molten salt electrolyte comprises a strong Lewis acid, and

5 electrochemically reducing the dissolved titanium oxide to titanium metal at a cathode.

In an embodiment of the invention there is provided a method for the direct production of titanium metal in a particulate state which comprises electrochemically dissolving an anode, formed of a titanium oxide/carbon composite, a molten salt

10 electrolyte in an electrochemical cell, wherein the molten salt electrolyte comprises a strong Lewis acid, and electrochemically reducing the dissolved titanium oxide to titanium metal.

Further features and advantages of the present invention will be seen by the following detailed description, taken in conjunction with the accompanying drawings

15 wherein:

Fig. 1 is a diagrammatic illustration schematically illustrating an electrochemical reaction according to the present invention;

Fig. 2a is a diagrammatic illustration of electrochemical process of the present invention;

20 Fig. 2b is a diagrammatic illustration of an electrochemical cell and process in accordance with the present invention;

Fig. 3 is a view similar to Fig. 2b providing further details of an electrochemical cell in accordance with the present invention;

25 Fig. 4 is a perspective view showing details of an electrode in accordance with the present invention;

Fig. 5 is a graph illustrating surface resistivity of a titanium oxide carbon anode over time.

The present invention employs a novel electrochemical system for producing titanium and other metals by a combination of thermal and electrochemical processes  
30 from a novel metal oxide-carbon composite anode. More particularly, the present invention produces purified titanium or other metal powders by a thermal/electrodeposition composite anode process using a metal oxide-carbon anode in a molten salt electrolyte.

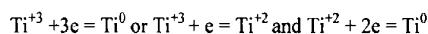
- 4a -

Heretofore the electrolysis of titanium oxide ( $TiO_2$ ) has not been successful because  $TiO_2$  has little to no solubility in molten salt electrolytes which is also true of other titanium compounds. Titanium tetrachloride ( $TiCl_4$ ) is a covalent compound that has limited solubility in fused salts and does not readily form complexes with

- 5 other inorganic salts. It also is highly volatile and is quickly lost from most fused salts. However, since titanium is multivalent, it has been shown that  $TiCl_4$  could be reduced to lower valent ionic species of  $Ti^{+3}$  and  $Ti^{+2}$ , which do exhibit some solubility in some molten salts. However, because of secondary reversibility reactions, which lead to loss in current efficiency and poor quality of metal,
- 10 heretofore no practical process has evolved for electrowinning titanium from a  $TiCl_4$  feed. Investigations of separating the anolyte and catholyte to avoid alternating oxidation and reduction with low current efficiency have not proven successful on a commercial scale.

Since titanium +3 (corresponding to  $y/x$  of 1.5) and titanium +2

- 15 (corresponding to  $y/x$  of 1.0) are ionic species, it should be possible to deposit titanium at the cathode, i.e. according to the reactions:



from a molten salt electrolyte. However, such reactions have not been demonstrated commercially since heretofore there has not been demonstrated an acceptable process

- 20 to continuously supply  $Ti^{+2y/x}$  or lower valence ions where  $y/x$  is less than 2 to a molten salt electrolyte. The present invention in one aspect provides a metal oxide/carbon composite anode containing  $Ti_xO_y$  in which a high valence metal such as  $Ti^{+4}$ , is thermally reduced to a valence less than +4, and is used to provide a continuous supply of reduced titanium ions to a molten salt electrolyte. The oxygen
- 25 combines with the carbon in the anode to produce  $CO_2$  and/or  $CO$  gas. Any excess carbon in the anode floats to the top of the molten salt electrolyte where it periodically

1 can be skinned if necessary and does not interfere with the continuous electrolysis  
2 process.

3 It is well established that thermal reduction is much more economical than  
4 electrochemical reduction. Therefore reducing  $TiO_2$  thermally is more economical  
5 than electrolytically reducing in a composite anode of  $TiO_2$ -carbon. If  $TiO_2$  is heated  
6 with carbon, carbo-thermic reduction will proceed based on the thermodynamic  
7 prediction and kinetics of the reactants. For example it has been found when the  
8 proper proportions of  $TiO_2$  and carbon are heated to various temperatures, reduced  
9 oxides are produced. An example reaction is  $2TiO_2 + C = Ti_2O_3 + CO$ . The  $Ti_2O_3$  in  
10 which the titanium is in a +3 valence state can be produced over the temperature  
11 range of 1250-1700°C. Since the product is a solid  $Ti_2O_3$  and gaseous CO if the  
12 pressure is reduced the kinetics of the reactions is enhanced.

13 It is also possible to produce the suboxide  $TiO$  according to the reactions  $TiO_2$   
14 + C =  $TiO$  + CO or  $Ti_2O_3$  + C =  $2TiO$  + CO. Either reaction will be enhanced at reduced  
15 pressure.

16 Titanium in  $TiO$  is in the +2 valence state. A competing reaction is  $TiO_2 + 3C = TiC$   
17 + 2CO or  $Ti_2O_3 + 5C = 2TiC + 3CO$ . When the suboxide is used as a feed for the  
18 composite anode, the lowest valence is the most desirable. Thus it is desirable to  
19 prevent  $TiC$  forming in which the titanium is in a +4 state. It has been found that  $TiO$   
20 can be produced at a reaction temperature above 1700°C if the pressure is reduced to  
21 0.01 atmosphere or lower. If the pressure is as high as 0.1 atmosphere a reaction  
22 temperature above 1800°C is required to produce  $TiO$  free of  $TiC$ . At atmospheric  
23 pressure a reaction temperature above 2000°C is required to produce  $TiO$  free of  $TiC$ .

24 In addition to producing titanium from a composite anode consisting of a  
25 reduced titanium oxide and a carbon source referred to as a composite anode it is also  
26 possible to electrowin titanium from other titanium compounds that are not oxides.  
27 These compounds include titanium nitride ( $TiN$ ). Titanium nitride is a conductor and  
28 does not require any conductive phase such as carbon with titanium suboxides.  $TiN$   
29 can be produced by reacting  $TiO_2 + 2C + N = TiN + 2CO$ . The  $TiN$  is pressed and  
30 sintered in a nitrogen atmosphere to produce a solid of  $TiN$ . The  $TiN$  can then be  
31 utilized as an anode in a fused salt to electrowin/deposit titanium at the cathode and  
32 nitrogen gas will be evolved at the anode.

33 Another compound is titanium carbide ( $TiC$ ). Titanium carbide is produced

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by the reaction of  $TiO_2 + 2C = TiC + 2CO$ . The TiC is a conductor and when TiC particles are pressed and sintered to a solid, the solid can serve as an anode. When using TiC as the anode a separator or diaphragm should separate the cathode and anode compartments. Titanium ions will be electrolytically dissolved from the anode and reduced to titanium metal at the cathode. The released carbon will be in solid form and must be accounted for in an overall materials balance. To account for the carbon the anode can be depolarized with oxygen wherein the oxygen will react with the carbon to form gaseous  $CO_2$  and/or CO. Thus oxygen gas would be passed over the anode to react with the carbon, but since titanium is so sensitive to oxygen the cathode should be separated from the anode with a diaphragm to prevent the oxygen from contacting the deposited titanium.

It is taught in WO09964638, US6,663,763B2, WO 02/066711 Al, WO 02/083993 Al and WO03/002 785 Al, that  $TiO_2$  can serve as a cathode in a calcium chloride fused salt wherein the  $TiO_2$  is reduced to titanium metal with oxygen given off at the anode using an inert anode or  $CO_2/CO$  using a carbon/graphite anode. Those teachings do not consider reduced or suboxides of titanium which require less electrochemical energy to produce titanium metal than required to reduce  $TiO_2$ . Thus the reduced oxides of  $Ti_2O_3$  or  $TiO$  can serve as cathodes and be electrochemically reduced in molten calcium chloride or other molten salt electrolytes.

Heretofore, there has not been an electrochemical system to produce titanium similar to electrowinning aluminum in which alumina ( $Al_2O_3$ ) is soluble in molten cryolite ( $NaAlF_4$ ) which under electrolysis produces aluminum metal with  $CO_2/CO$  being given off at a carbon anode, because there has not been identified a molten salt composition that will dissolve  $TiO_2$ . There is no known molten salt compound or combination of compounds that will dissolve  $TiO_2$ . However, there are molten salt compositions that will dissolve the reduced suboxide  $TiO$  which is an ionic compound that is very electrically conductive. For example  $TiO$  is soluble in molten calcium chloride mixed alkali and alkaline earth chlorides as well as fluorides or mixed chlorides and fluorides. Thus  $TiO$  can be dissolved in  $CaCl_2$  or other salt mixture, and using a carbon/graphite anode electrolyzed to produce titanium at the cathode and  $CO_2/CO$  at the anode or oxygen using an inert anode. Since titanium is sensitive to oxygen a separator or diaphragm should be used between the anode and cathode.

It is well known that the higher the temperature of a solvent the greater the

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solubility of the solute. In this case the higher the molten salt temperature the greater the solubility of a titanium suboxide such as  $TiO$  or  $Ti_2O_3$ . In the previous discussions the operating salt temperatures are below that of the melting point of titanium and thus titanium is deposited as a solid in a particulate morphology. As in the case of

- 5    electrowinning aluminum in which aluminum oxide is soluble in cryolite at over 900°C, the aluminum is in a molten state and thus more easily separated from the molten salt/cryolite. In order to achieve the same advantage with titanium, the molten salt operating temperature should be above the melting point of titanium or about 1670°C. Molten salts that have high melting temperatures that will not readily
- 10   vaporize at 1670°C or slightly above include calcium fluoride ( $CaF_2$ ) 1360°C, and barium fluoride  $BaF_2$  1280°C. It was found the titanium suboxides and particularly  $TiO$  is quite soluble in  $CaF_2$  at temperatures above 1670°C. Thus titanium is readily electrowon from its suboxides dissolved in  $CaF_2$  or other salts above 1670°C using a carbon/graphite anode that produces CO and  $CO_2$  on electrolysis or an oxygen stable
- 15   anode that produces oxygen on electrolysis. The titanium produced above 1670°C is in a molten state and thus readily separable from the molten salt whose density is less than 3.0 g/cc at the operating temperature and titanium is approximately 4.0 g/cc at the operating temperature thus causing the titanium to sink for easy separation.

Referring to Fig. 1, there is illustrated schematically the formation of a metal oxide-carbon composite anode in accordance with the present invention. Titanium oxide in a particle size of 0.001 - 1000 microns, preferably 0.01 - 500 microns, more preferably 0.1 to 10 microns, is mixed with carbon flakes of average particle size 0.001 - 1000 microns, preferably 0.01 - 100 microns, more preferably 0.01 to 1 microns, in a weight ratio of  $TiO_2$  to carbon of 7:1 to 4:1 using a ball mill mixer. The

- 25    $TiO_2$  powder and carbon flakes were mixed dry, or optionally with a binder, in a ball mill mixer for 4-24 hours. The resulting  $TiO_2$  powder/carbon flake mix was pressed in a steel die to form a mechanically stable green electrode or billet. The billet was then placed in an oven, and heated in the absence of air to 1000 to 2200°C, preferably about 1100°C to 1800°C, for 0.1 to 100 hours, preferably about two hours, to form a
- 30   titanium suboxide/carbon composite electrode.

Referring to Figs. 2a and 2b, the titanium oxide/carbon composite electrode 20 made as above described is employed as an anode in an electrochemical cell 22 with a

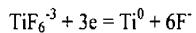
1 conventional metallic, e.g., steel electrode 24, and an alkali metal molten salt  
2 electrolyte 26.

3 The composition of the molten salt electrolyte 26 used in the cell 22 has an  
4 effect on the titanium produced at the cathode. The electrolyte should comprise a  
5 strong Lewis acid formulation such as NaAlCl<sub>4</sub>, which melts as low as 150°C,  
6 optionally containing fluoride additions such as an alkali fluoride and/or potassium  
7 titanium fluoride with the reduced state Ti<sub>x</sub>O<sub>y</sub>-C anode. Other useful electrolyte  
8 compositions include binary, tertiary, and quaternary alkali and alkaline earth  
9 chlorides, fluorides and mixed chloride-fluorides with melting point temperatures in  
10 the 300-900°C range. For producing titanium preferred electrolytes include NaCl -  
11 CaCl<sub>2</sub>-KCl in a mole ratio of 50:50:20; NaCl-LiCl-KCl in a mole ratio of 20:60:40;  
12 AlCl<sub>3</sub> - NaCl - NaF in a mole ratio of 70:30:20 L:Cl-KCl eutectic with 20 wt% NaF,  
13 eutectic of LiF-KF, etc. Moreover, the polarizing strength of the cation will directly  
14 affect the electroreduction of electrocrystallization to titanium. And, the small highly  
15 ionic strength and steric effect of e.g., a lithium ion in the electrolyte enhances the  
16 polarizing strength at the cathode and thus the electroreduction of titanium. Other  
17 such highly ionic ions can aid in stabilizing the Ti<sup>+3</sup> and/or Ti<sup>+2</sup> ions in the molten salt  
18 electrolyte as well as their electroreduction at the cathode.

19 To avoid disproportionation during the electrolysis between titanium in the  
20 metallic state, that is electrowon titanium, and higher titanium ions such as Ti<sup>+3</sup>, it is  
21 preferable to have only Ti<sup>+2</sup> ions in solution which as they are reduced to the metal are  
22 replaced with other Ti<sup>+2</sup> ions from the anode thus requiring TiO in the anode. Thus  
23 desirably the fused salt initially contains Ti<sup>+2</sup> ions which desirably is in the  
24 concentration range of ½ to 20%, more desirably in range of 1 to 10% and most  
25 desirably in the range of 2 to 8%.

26 The anion also can have an influence on the steric and solvent effect of the  
27 titanium species, which also influences the titanium deposit at the cathode. For  
28 example, the Ti-F bond is stronger than the Ti-C1 bond, which brings about an  
29 increase in the activity of the titanium ions in the molten salt electrolyte and  
30 consequently the morphology of the titanium deposited at the cathode. The anion and  
31 the titanium ion complex effects the number of crystallization centers available on the  
32 cathode and thus the morphology of the titanium cathode deposit. The complex TiF<sub>6</sub><sup>-3</sup>  
33 and the TiF<sub>6</sub><sup>-2</sup> anion is known and can be directly reduced to titanium. Mixed anions

are also known, such as  $\text{TiF}_{6-\text{N}} \text{Cl}_{\text{N}}^{-3}$ . A strong Lewis acid thus stabilizes and increases the activity of the titanium ion. While not wishing to be bound by theory, it is believed that the reactions proceed as follows:



5 and at the anode  $\text{Ti}^{+3}$  ions are released from the composite anode to produce the  $\text{TiF}_6^{-3}$ . Thus titanium is directly reduced from the +3 valence to the metal. Because titanium is multivalent it is also possible that  $\text{Ti}^{+3}$  is reduced to  $\text{Ti}^{+2}$  and then to the metal  $\text{Ti}^0$ . However, as stated above, if all titanium ions in solution are in the +2 valence then the reduction is  $\text{Ti}^{+2} + 2\text{e} = \text{Ti}^0$ .

10 Based on this analysis alkali fluorides may be regarded as stabilizing agents in chloride molten salt electrolytes. Thus the ratio of F/Cl and/or Ti/F will have an effect on the electroreduction of titanium. Indeed it has been demonstrated that all chloride molten salt electrolytes produce small and/or dendritic deposits of titanium. As fluorides are added to the molten salt electrolyte the morphology of the deposit

15 changes to larger and coherent particulate deposits. As the electrolyte changes to primarily or all fluoride, the titanium deposits become flaky to a fully adherent film. The major morphology change begins at a F/Cl ratio of approximately 0.1 and solid films become possible at a ratio of approximately 1.0.

The morphology and size of the titanium deposit is also influenced by the

20 current density of the cathode. The higher the current density the smaller the particle size. Typical cathode current densities are in the 0.05 to 5 amperes/cm<sup>2</sup> range. The more desirable cathode current densities are in the 0.1 to 2.0 amperes/cm<sup>2</sup> range, and the preferred cathode current densities are in the 0.25 to 1 amperes/cm<sup>2</sup> range, depending on the morphology of the titanium desired at the cathode. It also has been

25 found that very high current densities can be used at the cathode under high mass flow of the electrolyte and the use of the composite anode. By moving the electrolyte over the cathode surface via gas bubbling or pumping at a fast rate it is possible to electrolytically produce titanium particulate up to cathode current densities of 125 amps/cm<sup>2</sup>.

30 It also has been found that pulsing the current affects the morphology, particle size and cathodic efficiency. The current can be pulsed to on and off sequences in various wave forms such as square, sinusoidal, etc. as well as periodically alternating the polarity. It was found pulsing the current produced more coherent deposits and

larger particles as well as solid films on the cathode. It was also found periodically reversing the polarity between two composite electrodes produced titanium within the electrode. That is the  $Ti_xO_y$  in the electrode was reduced to titanium, which remained as a solid agglomerate of titanium particles in the same form of the original composite electrode.

5 A bench scale electrolytic cell for producing titanium in accordance with the present invention is illustrated in Fig. 3. The cell 30 comprises a cylindrically shaped steel walled vessel 32 having a funnel-shaped bottom closed by a valve 36. The vessel walls 32 are wrapped in a resistance heater (not shown) which in turn is covered by  
10 thermal insulation 40. A porous basket 42 formed of carbon fiber mesh is suspended within container 30 and is connected via an anode connector 44 to the plus side of the DC current source. Wall 32 of the steel vessel is connected via a conductor 46 to the negative side of a DC current source. Basket 42 is loaded with pellets or discs 48 of titanium suboxide - carbon flake anode material made as above described. The cell is  
15 filled with a molten salt electrode (60:LiCl - 40KCl) the cell is sealed with a top 50, swept with argon purge to remove air, and a voltage of 3.0V applied which resulted in precipitation of dendritic titanium sponge particles. The titanium sponge particles were then removed via valve 36, separated from the electrolyte, and found to have a purity of 99.9%.

20 It is possible to deposit other metals similarly. For example, by using a composite anode which includes other metal oxides in addition to the  $Ti_xO_y$ , it is possible to produce an alloy of titanium. For example, an alloy of Ti-Al-V can be produced by mixing aluminum oxide, vanadium oxide and  $TiO_2$  with carbon to form the anode whereby to produce alloy particulate or solid films of Ti-Al-V. The  $E_0$  and  
25 current density should be adjusted to deposit precise composition alloy particles. Other metals or alloys can be produced by incorporating other metal oxides in the anode in accordance with the present invention.

30 From a practical commercial standpoint of producing titanium particulate in which the particulate can be used directly in powdered metallurgical fabrication or consolidated into billets for subsequent rolling into sheet, forging, etc. it is desirable to produce the particulate at as low a cost as possible. High mass transfer and high current density that produces particle sizes that are desirable for commercial applications can be achieved in a cell configuration such as shown in Figure 4.

1           In this case the anode container can be a porous carbon-carbon or other anodic  
2    container in which  $Ti_xO_y$ -C anode segments 60 are placed, and the structural container  
3    can be the cathode and/or a cathode 62 placed inside the structural container (not  
4    shown). Preferably the container is insulated to maintain heat in the molten salt  
5    electrolyte to achieve thermal neutrality with the IR/joule heating of the electrolyte at  
6    high current densities. Also if desired the molten salt electrolyte could be pumped  
7    through cyclone systems and filters to continuously collect the titanium particulate as  
8    it is being produced. Commercial pumping systems are readily available to handle  
9    pumping molten salt electrolytes such as used in the aluminum and mass soldering  
10   industries to pump molten metals. Molten salt electrolytes that are desirable for high  
11   mass transfer cell designs of which Figure 4 is just one example, include strong Lewis  
12   acid compositions such as  $NaAlCl_4$  and fluoride compositions, and fluoride or  
13   chloride alkali and alkaline earth metal salts and mixtures thereof. Utilizing a high  
14   mass transfer cell design in which the molten salt electrolyte is pumped over the  
15   cathode surface, with high stirring rates and/or ultrasonics to agitate the molten salt  
16   electrolyte or the cathode itself coupled with a reduced valence  $Ti_xO_y$ -C anode  
17   permits production of titanium particulate at a relatively high rate and relatively low  
18   cost. And current pulsing as well as periodic reversing the current can further  
19   enhance the production of titanium particulates when coupled with a high mass  
20   transfer rate cell as above described.

21           Heretofore aluminum and magnesium have been produced by a composite  
22    anode process utilizing anodes of  $Al_2O_3$ -C or  $MgO$ -C [23-26]. However, there is no  
23    teaching a suggestion in any of the prior art that recognizes that high valence (4 or  
24    more) or multi-valence metals could be produced by a composite anode process.  
25    More importantly, it was not recognized that high value high valence or multi-valence  
26    metals such as titanium, chromium, hafnium, molybdenum, niobium, tantalum,  
27    tungsten, vanadium, and zirconium could be produced utilizing a composite anode, as  
28    in the present invention. Neither was it recognized that a high valence metal oxide  
29    could be thermally reduced to a lower valence state in a composite anode or that a  
30    reduced valence state metal oxide-carbon anode could be used to produce particulate  
31    metal by electroreduction.

32           In contrast to producing a molten metal aluminum (melting point approx. at  
33    660°C) and magnesium (melting point approx. at 650°C), the present invention

1 permits control of particle geometry and size, and grain size in the particle can be  
2 controlled by the molten salt composition, its operating temperature and the cathode  
3 current density. Moreover, the instant invention permits direct production of metals  
4 in the powdered/particulate solid state, unlike the prior art processes which produced  
5 molten aluminum [23, 25, 26] or magnesium [24].

6 In addition, the combination of thermal treatment to reduce the metal to a  
7 lower valence state, the use of carbon in the anode to release a lower valence state  
8 metal into the molten salt, and the selection of molten salt to stabilize the lower  
9 valence state metal so as to produce a fully reduced metal at the cathode, is a unique  
10 and advantageous feature of the current invention.

11 An alternative to reducing the titanium valence in the molten salt is to  
12 depolarize the cathode using hydrogen which could not only prevent the re-oxidation  
13 of the lower valence titanium at the anode and reduce the total cell voltage, but also  
14 allow for the formation of titanium hydrides at the cathodes. Titanium hydride is  
15 much more stable than titanium toward oxidation. The present invention thus permits  
16 the production of very low oxygen titanium.

17 Moreover, the present invention overcomes a problem of poor electrical  
18 conductivity of the metal oxide-carbon anode of my previous composite anode  
19 process [23-26] which required the use of aluminum or magnesium metal conductors  
20 through the composite anode to carry current and prevent high voltage drops due to  
21 the poor electrical conductivity of the Al<sub>2</sub>O<sub>3</sub>-C or MgO-C composite anodes. In the  
22 instant invention, poor anode electrical conductivity is overcome by using highly  
23 electrically conductive carbon flake as the major carbon source in the composite  
24 anode. Small size composite anode pieces can also be utilized to reduce voltage drop  
25 as illustrated in Figure 3 as contrasted to large size anodes which can result in high  
26 resistivity and larger voltage drops that increase energy consumption. Examples of  
27 low resistivity in a reduced valence state titanium oxide carbon anode is shown in  
28 Figure 5. Further when the TiO<sub>2</sub> is reduced to TiO, the TiO is very electrically  
29 conductive, more so than graphite. Thus anodes made with TiO are quite conductive  
30 and in one iteration does not require pressing into a composite with graphite flake or  
31 other carbon forms. The TiO is so conductive, it can be simply mixed with  
32 carbon/graphite in a basket that serves as the anode with a conductor which can be the  
33 basket or a graphite rod.

1           The following non-limiting Examples will further demonstrate the present  
2   invention.

3           Example 1

4           Titanium dioxide ( $TiO_2$ ) with a purity of 99% in a particle size of 0.3 microns  
5   was mixed with graphite flake in a particle size of 40 microns in a ratio of 80 grams of  
6    $TiO_2$  and 20 grams of graphite flake using a ball mill mixer. The resulting  $TiO_2$ -  
7   graphite flake mixture was pressed in a steel die at 50,000 psi, which provided a  
8   mechanically stable billet without any binder system. The  $TiO_2$ -graphite flake billet  
9   was heated to 1100°C in the absence of air for two hours. An XRD analysis showed  
10   the resulting composite anode to consist of  $Ti_2O_3$ ,  $Ti_3O_5$  and  $Ti_4O_7$  and graphite. The  
11   resulting titanium oxide-graphite composite anode was cut into one inch (2.54 cm)  
12   long segments, and the segments placed in a carbon-carbon composite basket as  
13   illustrated in Fig. 3 which had residual porosity that served as a membrane and to  
14   which the positive terminal of a dc power supply was connected. A steel walled  
15   container (illustrated in Fig. 3) was used to melt an electrolyte consisting of  $NaCl$ -  
16    $CaCl_2$ - $KCl$  eutectic at a temperature of 650°C. The steel walled container was  
17   connected to the negative terminal of the dc power supply. The steel walled container  
18   was covered, sealed and swept with an argon purge to remove any air from the  
19   system. Electrolysis was conducted at an anode and cathode current density of 0.5  
20   amps/cm<sup>2</sup>, which produced titanium particulate at the steel cathode. The titanium  
21   particulate was harvested with a screen scoop and then subjected to 1200°C in a  
22   vacuum to remove all traces of the electrolyte. The particle size was in the range of  
23   one to ten microns with a predominance of 5-10 microns. The titanium powder was  
24   analyzed for oxygen and found to have 800 parts per million. The current efficiency  
25   was measured by calculating the ampere hours passed and weighing the titanium  
26   produced which was found to be 95% at the cathode and 99% at the anode.

27           Example 2

28           A mixture of  $TiO_2$  and graphite flake was mixed as described in Example 1  
29   and a resin binder of phenolic was used to bind the particles in the pressing operation.  
30   The pressed body was then heated in an inert atmosphere to 1300°C, which produced  
31   a well-bonded strong composite anode consisting of a mixture of  $Ti_2O_3$  with some  
32    $Ti_3O_5$  and a small amount of  $TiC$  along with graphite. Electrolysis was conducted as

1 in Example 1 at a cathode current density of 1.0 amp/cm<sup>2</sup>. Titanium powder was  
2 produced at an efficiency of 90% in an average particle size of 10 microns.

3 Example 3

4 Example 2 was repeated with electrolysis at a cathode current density of 0.25  
5 amps/cm<sup>2</sup> which produced an efficiency of 97% with a particle size of approximately  
6 20 microns.

7 Example 4

8 A composite anode was produced using a mixture of TiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub> and V<sub>2</sub>O<sub>3</sub> in  
9 an elemental ratio of Ti-6Al-4V. A stoichiometric ratio of graphite flake was mixed  
10 with the oxides and a coal tar pitch binder was used. The molded composite anode  
11 was heat treated to 1200°C in an inert atmosphere. The composite anode was placed  
12 in the anode basket as described in Example 1 but a sheet of titanium was used as the  
13 cathode. The electrolyte consisted of NaCl-LiCl-KCl eutectic with 20 mole % LiF.  
14 Electrolysis was conducted at a cathode current density of 1.25 amps/cm<sup>2</sup>, which  
15 produced particles in a size primarily in the range of 10-80 microns. The harvested  
16 particles were analyzed and found to contain a ratio of Ti-6Al-4V.

17 Example 5

18 A composite anode was prepared as described in Example 1 and heat treated  
19 to 1150°C. The molten salt electrolyte consisted of KF-NaF-LiF eutectic operated at  
20 650°C. The cathode was nickel metal with electrolysis conducted at a cathode current  
21 density of 0.25 amps/cm<sup>2</sup>. A coherent film of titanium 10 microns thick was  
22 deposited on the nickel cathode.

23 Example 6

24 A composite anode was produced as described in Example 2 using Y<sub>2</sub>O<sub>3</sub> and  
25 graphite flake in stoichiometric ratio. The anode was electrolyzed as in Example 2,  
26 which produced yttrium metal in a particle size of 10-30 microns.

27 Example 7

28 A composite anode was produced as described in Example 2 using  
29 stoichiometric ratio of HfO<sub>2</sub> and carbon. Electrolysis of the anode in a molten salt  
30 electrolyte, as in Example 4, at a cathode current density of 0.5 amperes/cm<sup>2</sup>  
31 produced metal hafnium metal particulate having a particle size of 10 - 100 microns.

32

### Example 9

Rutile ore which contained approximately 95% TiO<sub>2</sub> was dried and mixed with graphite flake and a resin binder to produce the oxide-carbon in stoichiometric ratio. The mixture was compressed to 20,000 psi and heat treated in an inert atmosphere to 1200°C. The anode was electrolyzed as in Example 4, which produced a powder at the cathode containing primarily titanium, and small amounts of iron, aluminum, niobium, vanadium and silicon having a particle size of 1 - 80 microns.

### Example 10

15 A salt composition of (65 AlCl<sub>3</sub> – 35 NaCl mole %) -20 mole % NaF was  
16 utilized as the electrolyte at an operating temperature of 190°C. A composite anode  
17 was utilized as described in Example 1 with electrolysis conducted with a pulsed  
18 current 3 seconds on and 1 second off. A crystalline titanium deposit of flake  
19 morphology was produced at a cathode current density of 1 amps/cm<sup>2</sup>.

### Example 11

21 Example 10 was repeated with a cathode current density of 0.25 amps/cm<sup>2</sup>.  
22 The resulting titanium deposit was a solid film on the cathode. The pulse scheme was  
23 then modified to 3 seconds on ¼ second off with periodic reverse polarity and then  
24 repeating the cycle. The deposit was a solid film with a very fine grain  
25 microstructure. Other shape form pulses provided similar results.

### Example 12

27 Hydrogen was used at the cathode in an electrolytic cell similar to Example 10  
28 with or without a pulsed current. Cell voltage was decreased by about 10 to 15%, and  
29 titanium hydride powder formed in-situ in the cell instead of metallic titanium  
30 powder. Washing the titanium hydride produced oxygen pick up of  $\leq 200$  ppm. The  
31 resulting titanium hydride was then dehydrogenated by heating to about 650°C to  
32 produce metallic titanium powder with  $\leq 400$  ppm oxygen. This oxygen level is an  
33 order of magnitude lower than titanium powder produced by any other process.

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Example 13

Titanium oxide was mixed with a stoichiometric amount of carbon black and heated under a reduced pressure of 0.01 atmosphere in argon to a temperature of 1450<sup>0</sup>C which produced the titanium suboxide of Ti<sub>2</sub>O<sub>3</sub> with no other suboxides or 5 contaminates such as TiC. The Ti<sub>2</sub>O<sub>3</sub> was mixed with graphite flake, a binder of phenolic resin, and pressed into a block. The block was heated in the absence of air to 1100<sup>0</sup>C which formed an anode. The resulting composite anode was used in a fused salt consisting of the eutectic of LiCl-KCl operated at 500<sup>0</sup>C. Electrolysis was conducted in trial one at 1 amp/cm<sup>2</sup> on the cathode which produced titanium 10 particulate in a size of 1 to 10 microns. In a second trial a titanium sponge was placed in the bottom of the fused salt and TiCl<sub>4</sub> was bubbled onto the sponge which produced TiCl<sub>2</sub> in the salt bath. TiCl<sub>4</sub> continued until a concentration of 5% TiCl<sub>2</sub> was generated. Electrolysis was then performed as in trial one and titanium particulate 15 with a size up to 400 microns was produced, thus showing with a titanium ion in solution larger size titanium particulate was produced.

Example 14

An identical system as in Example 13 was created before and TiCl<sub>2</sub> was generated, and in trial one the electrolysis was performed at 40 amps/cm<sup>2</sup>. The titanium particulate produced was in a size range of 20 to 100 microns. In trial two, 20 electrolysis was performed at 125 amps/cm<sup>2</sup> which produced titanium particles in approximately the same size as the 40 amps/cm<sup>2</sup> current density trial. In trial three, electrolysis was also performed at 125 amps/cm<sup>2</sup> with argon gas bubbling over the cathode to create a large mass flow. The titanium particulate produced in the high mass flow at 125 amps/cm<sup>2</sup> was in the size range of 40 to 200 microns. The titanium 25 suboxide-carbon composite anode provides the opportunity to operate at very high cathode current densities and in a high mass flow regime.

Example 15

TiO<sub>2</sub> and carbon were heated under a pressure of 0.01 residual argon atmosphere to 1850<sup>0</sup>C which produced TiO and CO. The TiO was mixed with 30 stoichiometric carbon and a binder and molded into a block which was heated to 1100<sup>0</sup>C which formed a composite anode. The resulting composite anode was placed in a salt mixture of 60NaCl-40MgCl<sub>2</sub> and 20 mole percent NaF based on the chloride salt mixture operated at 600<sup>0</sup>C. In trial one, the electrolysis was performed at 0.15

1     amps/cm<sup>2</sup> and titanium particulate in the size range of 50 to 300 microns was  
2     produced. In trial two, a titanium sponge was placed in a small crucible immersed in  
3     the salt bath and TiCl<sub>4</sub> was bubbled onto the sponge that produced TiCl<sub>2</sub> until the  
4     concentration was 8% TiCl<sub>2</sub> in the salt. Electrolysis was performed at 0.15 amps/cm<sup>2</sup>  
5     which produced titanium particulate in the 200 to 500 micron size. The oxygen  
6     content was 380 parts per million.

### Example 16

Rutile with a composition as follows, and the remainder titanium was processed as shown in the headings:

Impurities	Units	As received composition	After heating to 1700°C with carbon	Purity of Electrolytically produced titanium
Al	ppm	5300	4200	700
Ca	ppm	570	530	<100
Cr	ppm	300	150	100
Fe	ppm	4390	140	100
Mg	ppm	1470	1270	500
Si	ppm	12000	<100	<100
V	ppm	2290	2290	2000
Zr	ppm	360	250	300

With the remainder titanium  
10  
11 The rutile was mixed with carbon in a ratio of 1.1 to stoichiometry and heated to  
12 1700°C in argon at atmospheric pressure. The composition after heating is shown in  
13 the second column which shows the rutile was purified and particularly in the case of  
14 iron and silicon of which the latter is most undesirable as an impurity in titanium  
15 metal.

16 The purified rutile was mixed with carbon and resin and molded onto a block  
17 which was heat treated to 1250°C. The composite block was utilized as an anode in a  
18 salt bath of NaCl-CaCl<sub>2</sub> operated at 650°C. Electrolysis was performed at 0.5  
19 amps/cm<sup>2</sup> which produced particulate in the size range of 50-350 microns with a  
20 purity as shown in column five above. Aluminum and vanadium are desirable  
21 alloying elements for titanium and are used in most titanium alloys. Thus a relatively

1 pure titanium is produced from low cost domestic source rutile which can meet  
2 virtually all market demands except the stringent aerospace requirements.

3 Example 17

4  $\text{TiO}_2$  was mixed with carbon and heated in a 90% nitrogen 10% hydrogen  
5 atmosphere to  $1600^{\circ}\text{C}$  which produced titanium nitride (TiN). The TiN was pressed  
6 and sintered at  $2000^{\circ}\text{C}$  in a nitrogen atmosphere. The TiN block was used as an anode  
7 in a salt mixture of (NaCl-KCl) - 20 mole % NaF operated at  $700^{\circ}\text{C}$ . Electrolysis was  
8 conducted at 0.5 amps/cm<sup>2</sup> which produced titanium particulate in the size range of  
9 20 to 350 microns and nitrogen gas was given off at the anode.

10 Example 18

11  $\text{TiO}_2$  was mixed with carbon in a ratio of 1 to 1.5 over stoichiometry and  
12 heated in argon at  $1600^{\circ}\text{C}$  which produced titanium carbide (TiC). The TiC was  
13 pressed and sintered at  $2000^{\circ}\text{C}$ . The TiC was used as an anode in the same salt as in  
14 Example 17. During electrolysis at 1 amp/cm<sup>2</sup> oxygen was bubbled under the TiC  
15 anode in an amount equivalent to the current to produce titanium so that the oxygen  
16 reacted with the freed carbon to produce  $\text{CO}_2/\text{CO}$  which is often referred to as  
17 depolarizing the electrode. A diaphragm of porous alumina was placed between the  
18 anode and cathode to prevent any oxygen from contacting the deposited titanium  
19 particulate and oxidizing it. The particle size of titanium particulate produced was  
20 in the size range of 20 to 200 microns.

21 Example 19

22 The suboxide  $\text{TiO}$  was produced by reacting  $\text{TiO}_2$  with stoichiometric carbon  
23 in a vacuum of 0.01 atmosphere at a temperature of  $1850^{\circ}\text{C}$ . The  $\text{TiO}$  was then  
24 pressed and practically sintered at  $1450^{\circ}\text{C}$  to provide a porous body which served  
25 as a cathode in a fused salt bath of calcium chloride containing 5% calcium oxide  
26 operated at  $900^{\circ}\text{C}$ . A graphite anode was utilized and electrolysis performed at a  
27 constant voltage of 3.0V for a period of 12 hours. The  $\text{TiO}$  was reduced to titanium  
28 metal with oxygen being attracted to the anode to produce  $\text{CO}_2/\text{CO}$ .

29 Example 20

30 Example 19 was repeated using  $\text{Ti}_2\text{O}_3$  as the starting material.

31 Example 21

32 Example 19 was repeated with the exception the electrolyte was the eutectic  
33 of  $\text{CaCl}_2\text{-NaCl}$  which was operated at  $750^{\circ}\text{C}$ . With the suboxide  $\text{TiO}$ , the lower

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temperature salt bath can be used to reduce TiO to titanium metal.

Example 22

The molten salt bath electrolyte of  $\text{CaCl}_2$  operated at  $900^\circ\text{C}$  showed a considerable solubility of the reduced suboxide of titanium TiO. In a salt bath operated at  $900^\circ\text{C}$  5 wt % TiO was added and electrolysis conducted with a carbon anode. Titanium particulate was deposited on the cathode at a current density of  $1 \text{ amp/cm}^2$ . In a second trial, a porous alumina diaphragm was used around the anode to prevent any oxygen from diffusing to the deposited titanium on the cathode and contaminating the deposited titanium particulate.

10 Example 23

A molten salt composition consisting of the  $\text{CaCl}_2\text{-NaCl}$  eutectic containing 20 mole % NaF was operated at  $750^\circ\text{C}$  and 2 wt % TiO was added which became soluble in the salt bath. A carbon anode was used and electrolysis performed at a cathode current density of  $0.25 \text{ amps/cm}^2$ . Titanium particulate was deposited on the cathode and  $\text{CO}_2/\text{CO}$  was evolved from the carbon anode.

Example 24

TiO was produced as described in Example 15 and mixed with carbon particulate. The mixture of TiO-C was placed in a porous carbon-carbon basket which served as the anode electrical conductor. The anode basket containing TiO-C was placed in a salt of LiCl-KCl eutectic containing 20 wt% NaF operated at  $680^\circ\text{C}$ . Electrolysis was performed at  $1 \text{ amps/cm}^2$  which produced titanium particulate in the size range of 50-500 microns which demonstrated a physical mixture of TiO-C can serve as an anode.

Example 25

25 An anode produced as described in Example 13 was utilized in the electrolyte given in Example 13 with electrolysis conducted at  $1 \text{ amps/cm}^2$  concurrent with hydrogen bubbling under the cathode. The deposit was titanium particulate in the size range of 50-800 microns. Heating the deposit showed hydrogen evolution as detected in a mass spectrometer.

30 Example 26

A graphite crucible was set inside a steel cell with a cover and seal to provide an inert atmosphere with an argon purge. A graphite rod with a reduced tip to serve as a resistor was placed through a standard feed-through in the cell cover. Calcium

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- 20 -

fluoride was placed in the graphite crucible. The graphite rod was heated resistively between a connection to it and the steel cell which raised the temperature to 1700°C which melted the calcium fluoride. TiO was then added at 5 wt%. Electrolysis was conducted at 1 amps/cm<sup>2</sup> between a separate graphite anode and the crucible serving as the cathode. After six hours of electrolysis the experiment was stopped and the system cooled. Breaking the salt revealed beads of titanium that had been produced in the molten salt.

Example 27

Example 26 was repeated with a graphite resistor heater located between two graphite rods that melted the calcium fluoride and raised the temperature to 1710°C. Ti<sub>2</sub>O<sub>3</sub> was then added at 10wt% of the melted CaF<sub>2</sub>. Electrolysis was conducted between a tungsten cathode and a platinum-iridium anode at a current density of 0.5 amps/cm<sup>2</sup>. During the electrolysis oxygen was given off at the anode which acted as a non-consumable inert anode in contrast to graphite which forms CO and CO<sub>2</sub>. After five hours operation the experiment was stopped and the molten portion of the molten salt cracked which revealed numerous beads of titanium metal.

The above embodiments and examples are given to illustrate the scope and spirit of the instant invention. These embodiments and examples are within the contemplation of the present invention. Therefore, the present invention should be limited only by the appended claims.

Throughout this specification and the claims which follow, unless the context requires otherwise, the word "comprise", and variations such as "comprises" and "comprising", will be understood to imply the inclusion of a stated integer or step or group of integers or steps but not the exclusion of any other integer or step or group of integers or steps.

The reference to any prior art in this specification is not, and should not be taken as, an acknowledgment or any form or suggestion that the prior art forms part of the common general knowledge in Australia.

30

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The claims defining the invention are as follows:

1. A method for the production of titanium metal which comprises electrochemically dissolving, in a molten salt electrolyte, an anode formed of a titanium suboxide/carbon composite, and reducing the dissolved titanium suboxide, at a cathode, to titanium metal.
2. A method for the production of purified titanium from rutile ore which comprises electrowinning from an anode formed of a mixture of titanium suboxide/carbon composite in a molten salt electrolyte, and depositing purified titanium onto a cathode.
3. A method for the production of purified titanium which comprises electrochemically dissolving an anode formed of a titanium suboxide/carbon composite in a molten salt electrolyte, and electrochemically reducing the dissolved titanium suboxide to purified titanium metal.
4. A method for the direct production of titanium metal in a particulate state which comprises electrochemically dissolving an anode, formed of a titanium suboxide/carbon composite, in a molten salt electrolyte in an electrochemical cell, and electrochemically reducing the dissolved titanium suboxide to particulate titanium metal.
5. A method for the production of titanium metal which comprises electrochemically dissolving, in a molten salt electrolyte, an anode formed of a titanium oxide/carbon composite, wherein the molten salt electrolyte comprises a strong Lewis acid, and electrochemically reducing the dissolved titanium oxide to titanium metal at a cathode.
- 30 6. A method for the direct production of titanium metal in a particulate state which comprises electrochemically dissolving an anode, formed of a titanium oxide/carbon composite, a molten salt electrolyte in an electrochemical cell, wherein the molten salt electrolyte comprises a strong Lewis acid, and electrochemically

reducing the dissolved titanium oxide to titanium metal.

7. The method of claims 1 or 4, wherein said molten salt electrolyte comprises a strong Lewis acid.

5

8. The method of any one of claims 1, 2, 3, 4, 5 or 6, wherein the electrolyte is selected from the group consisting of an eutectic of sodium chloride, lithium chloride and potassium chloride, an eutectic of potassium fluoride, sodium fluoride and lithium fluoride, an eutectic of sodium chloride, calcium chloride and potassium chloride, an eutectic of sodium chloride, magnesium chloride and sodium fluoride, and an eutectic of sodium chloride, potassium chloride and sodium fluoride.

10

15

9. The method of claims 2 or 3, wherein titanium suboxide is mixed with carbon in a ratio of at least 1:1.5 over stoichiometry preferably in a ratio of at least 1:1 over stoichiometry based on titanium to produce TiC and CO<sub>2</sub>/CO.

10. The method of claim 3, wherein titanium suboxide-carbon composite anode is formed by heating a titanium oxide with carbon under an inert atmosphere.

20

11. The method according to claim 3, wherein the anode comprises a composite of titanium suboxide and carbon, and including the step of adding a Ti<sup>+2</sup> containing compound to the electrolyte.

25

12. The method of claim 4, wherein the electrolyte includes a Ti<sup>+3</sup> containing compound which is reduced in one step to titanium metal.

13. The method of claim 12, wherein the Ti<sup>+3</sup> containing compound is added in a concentration of 0.5 to 20% by weight of the electrolyte.

30

14. The method according to claims 4 or 5, wherein the electrode is formed of a titanium oxide/carbon composite, and including the step of adding a Ti<sup>+2</sup> containing compound to the electrolyte.

15. The method of claim 14, wherein the  $Ti^{+2}$  containing compound is added in a concentration of 0.5 to 20% by weight of the electrolyte.

5 16. The method of claim 6, wherein the electrolyte includes  $Ti^{+3}$  containing compound which is reduced in one step to titanium metal.

17. The method of claim 16, wherein the  $Ti^{+3}$  containing compound is added in a concentration of 0.5 to 20% by weight of electrolyte.

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18. A method for the production of titanium meal from its oxide ore, substantially as herein described with reference to Example 1 or Example 2.

19. The method of claim 13 or 17, wherein the  $Ti^{+3}$  containing compound is 15 added in a concentration of 1 to 10% by weight of the electrolyte.

20. The method of claim 15, wherein the  $Ti^{+2}$  containing compound is added in a concentration of 1 to 10% by weight of the electrolyte.

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FIG. 1

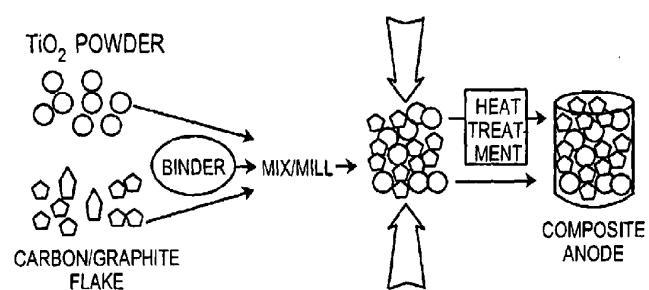
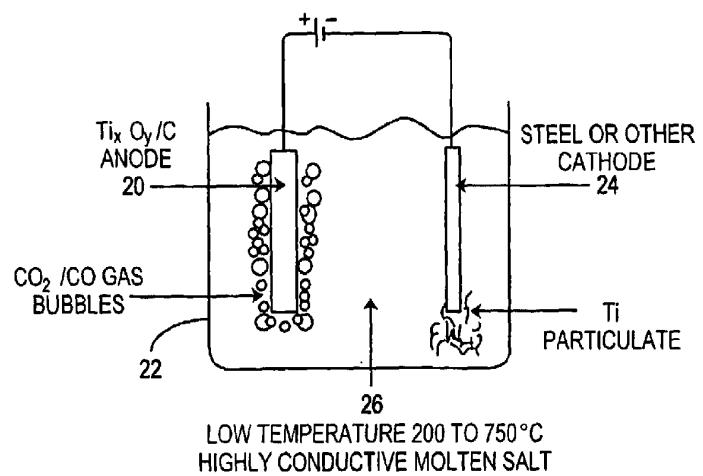


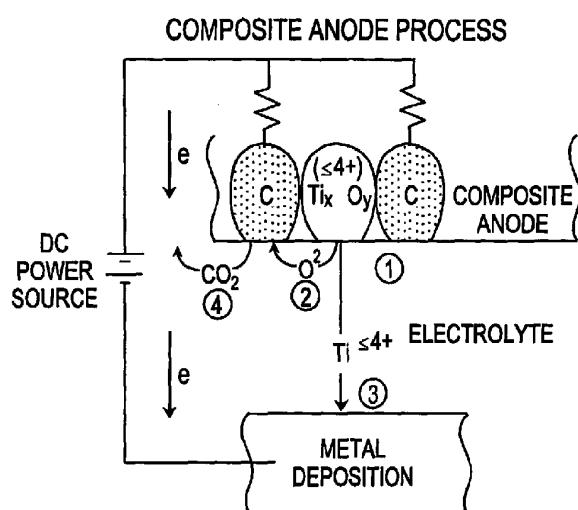
FIG. 2b



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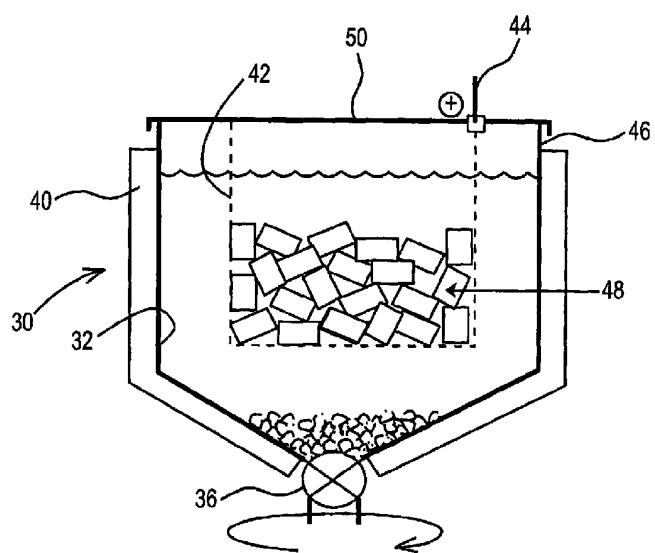
FIG. 2a



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FIG. 3



SUBSTITUTE SHEET (RULE 26)

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FIG. 4

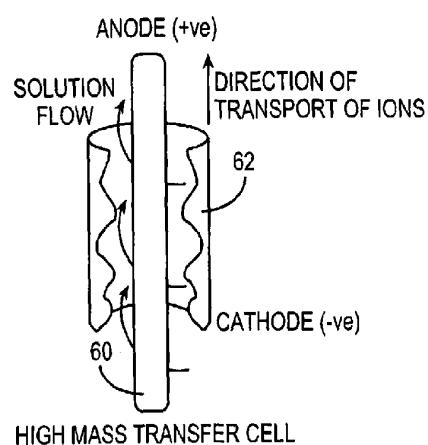


FIG. 5

