

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property

Organization

International Bureau

(43) International Publication Date

04 June 2020 (04.06.2020)



(10) International Publication Number

WO 2020/112899 A1

(51) International Patent Classification:

B01D 53/22 (2006.01) B01J 31/16 (2006.01)

B01J 20/22 (2006.01)

(21) International Application Number:

PCT/US2019/063442

(22) International Filing Date:

26 November 2019 (26.11.2019)

(25) Filing Language:

English

(26) Publication Language:

English

(30) Priority Data:

62/771,537 26 November 2018 (26.11.2018) US

(71) Applicant: **THE REGENTS OF THE UNIVERSITY OF CALIFORNIA** [US/US]; 1111 Franklin Street, 12th Floor, Oakland, California 94607 (US).

(72) Inventors: **YAGHI, Omar M.**; 2150 Shattuck Ave., Suite 510, Berkeley, California 94704 (US). **HANIKEL, Nikita**; 2150 Shattuck Ave., Suite 510, Berkeley, California 94704 (US). **LYU, Hao**; 2150 Shattuck Ave., Suite 510, Berkeley, California 94704 (US).

(74) Agent: **OSMAN, Richard Aron**; 530 Lawrence Expy #332, Sunnyvale, California 94085 (US).

(81) Designated States (*unless otherwise indicated, for every kind of national protection available*): AE, AG, AL, AM, AO, AT, AU, AZ, BA, BB, BG, BH, BN, BR, BW, BY, BZ, CA, CH, CL, CN, CO, CR, CU, CZ, DE, DJ, DK, DM, DO, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, GT, HN, HR, HU, ID, IL, IN, IR, IS, JO, JP, KE, KG, KH, KN, KP, KR, KW, KZ, LA, LC, LK, LR, LS, LU, LY, MA, MD, ME, MG, MK, MN, MW, MX, MY, MZ, NA, NG, NI, NO, NZ, OM, PA, PE, PG, PH, PL, PT, QA, RO, RS, RU, RW, SA, SC, SD, SE, SG, SK, SL, SM, ST, SV, SY, TH, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, ZA, ZM, ZW.

(84) Designated States (*unless otherwise indicated, for every kind of regional protection available*): ARIPO (BW, GH, GM, KE, LR, LS, MW, MZ, NA, RW, SD, SL, ST, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, RU, TJ, TM), European (AL, AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HR, HU, IE, IS, IT, LT, LU, LV, MC, MK, MT, NL, NO, PL, PT, RO, RS, SE, SI, SK, SM, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, KM, ML, MR, NE, SN, TD, TG).

Published:

— with international search report (Art. 21(3))

(54) Title: MULTIVARIATE AND OTHER METAL-ORGANIC FRAMEWORKS, AND USES THEREOF

(57) Abstract: Compositions comprising multivariate metal-organic frameworks and other single-linker metal-organic frameworks are used for containing and storing a gas or fluid or for water harvesting or water purification applications.



WO 2020/112899 A1

MULTIVARIATE AND OTHER METAL-ORGANIC FRAMEWORKS, AND USES
THEREOF

CROSS-REFERENCE TO RELATED APPLICATION

[0001] This application claims priority to U.S. Provisional Patent Application No. 62/771,537, filed November 26, 2018, which is incorporated herein by reference in its entirety.

FIELD

[0002] The present disclosure relates generally to water harvesting, and more specifically to systems and methods for harvesting water from surrounding air using various types of metal-organic frameworks.

BACKGROUND

[0003] Water scarcity has been indicated as one of the major evolving global risks by the World Economic Forum in the recent Global Risks Report 2018.¹ About two thirds of the global population faces water scarcity for at least one month of the year.² The water supply crisis is expected to exacerbate with growing population and ongoing climate change.³ Thus, different strategies were proposed and implemented to cope with this global challenge.⁴

[0004] One realm of possible solutions relies on the usage of atmospheric humidity as an ubiquitous water resource.⁵ This includes fog harvesting⁶, water collection via dewing⁷ or water vapor concentration by desiccants.⁸ However, these technologies require either the occurrence of a very high relative humidity ($RH \approx 100\%$) or high energetic investment for regeneration of the employed sorbent.⁹

[0005] Only recently, metal-organic frameworks (MOFs) were found to be a promising class of materials circumventing the above described challenges of water harvesting from air.⁹ In particular, MOFs exhibit desirable properties for water harvesting applications under arid conditions, *i.e.* adsorption of large amounts of atmospheric water at low RH and successive energy-efficient desorption without loss of porosity.

[0006] While the practical employment of MOFs for water harvesting has been realized recently,¹⁰ only few MOFs are reported to be suitable for water harvesting under arid conditions.¹¹ This invention provides a guideline to fine-tune and improve the water sorption properties by employing metal-organic frameworks constructed from a combination of two or more building blocks, which may generally be referred to as multivariate (MTV) MOFs.¹² This

approach will not only drastically expand the variety of water harvesting systems but will also allow for adjustment of the water sorption properties for the respective application.

BRIEF SUMMARY

[0007] The invention provides multivariate and other metal-organic frameworks (MOFs) for water harvesting and other uses. In some variations, the multivariate metal-organic frameworks (MTV-MOFs) may have mixed metals in the secondary building units and/or mixed organic ligands.

[0008] In some aspects, provided is a MOF comprising repeating cores, wherein the cores comprise secondary building units (SBUs) connected to organic ligands. In some variations, the secondary building units comprise one or more metals or metal-containing complexes. In one variation, the secondary building units form 1D rod-like chains or distinct multinuclear metal clusters. In certain variations, the organic ligands comprise one or more linear ditopic moieties, v-shaped ditopic moieties, trigonal tritopic moieties, square or rectangular tetratopic moieties, or tetrahedral tetratopic moieties. In one variation, such moieties comprise 5-membered or 6-membered rings substituted with at least two carboxylate groups.

[0009] In an aspect the invention provides a multivariate metal-organic framework (MTV-MOF) of inorganic metal clusters (secondary building units, SBUs) and a combination of two or more different organic units/ligands (linkers).

[0010] In another aspect, provided is a MOF made up of secondary building units and one type of organic ligand.

[0011] In some embodiments of the foregoing, the organic ligands are linear ditopic, v-shaped ditopic, trigonal tritopic, square or rectangular tetratopic, or tetrahedral tetratopic. In one variation, at least one of the organic ligands is v-shaped as disclosed herein.

[0012] In some variations, the secondary building units are metal clusters, or comprise a metal or metal-containing complex. In certain variations, the secondary building units form 1D rod-like chains or distinct multinuclear metal clusters.

[0013] In some embodiments, the metal in the secondary building unit is selected from zirconium, nickel, iron, copper, manganese and aluminum, and/or from magnesium, calcium, strontium, barium, titanium, zinc, indium, cadmium, hafnium, lead, cobalt, and chromium.

[0014] In an aspect the invention provides a device such as a water harvester, comprising a moisture sorption unit comprising one or more of the MOFs described herein.

[0015] In an aspect the invention provides a method of using one or more of the MOFs described herein, comprising containing in, storing in and/or extracting from the composition a predetermined gas or fluid, such as CO₂, H₂O, H₂, CH₄, C₂H₄, C₂H₂, etc.

[0016] In an aspect the invention provides a method of using one or more of the MOFs described herein, comprising utilizing the composition for water harvesting or water purification applications.

[0017] In another aspect the invention provides a water harvester comprising a moisture sorption unit comprising a MOF of formula Al(OH)(2,5-PyIDC), wherein 2,5-PyIDC is 2,5-pyrroledicarboxylate.

[0018] In an aspect the invention provides a method of using the water harvester comprising for water harvesting or water purification applications.

[0019] The invention encompasses all combination of the particular embodiments recited herein, as if each combination had been laboriously recited.

DESCRIPTION OF THE FIGURES

[0020] The present application can be understood by reference to the following description taken in conjunction with the accompanying figures.

[0021] FIGS. 1A and 1B depict a non-comprehensive list of compounds that may be used to form organic ligands connected to secondary building units in the MOFs described herein. These linkers can also be combined in systems for water harvesting to expand the realm of MOF materials for water harvesting.

[0022] FIGS. 2 and 3 depict a structural model of Al(OH)(2,5-PyIDC). FIG. 2 shows views along the a-axis and c-axis. FIG. 3 shows an excerpt of the secondary building units of the MOF. Oxygen atoms are located at the vertices of the octahedron, coordinating to an aluminum atom at the center of the octahedron.

[0023] FIG. 4 depicts a comparison of the powder x-ray diffraction (PXRD) data of the experimentally attained MOF Al(OH)(2,5-PyIDC) and its proposed structural model.

[0024] FIG. 5 depicts a nitrogen sorption analysis on Al(OH)(2,5-PyIDC) conducted at 77 K (BET surface area = 1051 m²/g).

[0025] FIG. 6 depicts a water sorption isotherm of Al(OH)(2,5-PyIDC) conducted at 25 °C.

[0026] FIG. 7 depicts a structural model of Al(OH)(2,4-PyIDC): view along the a-axis (left) and c-axis (right). Oxygen atoms are located at the vertices of the octahedron, coordinating to an aluminum atom at the center of the octahedron.

[0027] FIG. 8 depicts a comparison of the powder x-ray diffraction (PXRD) data of the experimentally attained MOF Al(OH)(2,4-PyIDC) and its proposed structural model.

[0028] FIG. 9 depicts a nitrogen sorption analysis on Al(OH)(2,4-PyIDC) conducted at 77 K (BET surface area = 912 m²/g).

[0029] FIG. 10 depicts a water sorption isotherm of Al(OH)(2,4-PyIDC) conducted at 25 °C.

[0030] FIG. 11 depicts a structural model of Al(OH)(2,4-TDC): view along the a-axis (left) and c-axis (right). Oxygen atoms are located at the vertices of the octahedron, coordinating to an aluminum atom at the center of the octahedron.

[0031] FIG. 12 depicts a comparison of the powder x-ray diffraction (PXRD) data of the experimentally attained MOF Al(OH)(2,4-TDC) and its proposed structural model.

[0032] FIG. 13 depicts a nitrogen sorption analysis on Al(OH)(2,4-TDC) conducted at 77 K (BET surface area = 1101 m²/g).

[0033] FIG. 14 depicts a water sorption isotherm of Al(OH)(2,4-TDC) conducted at 25 °C.

[0034] FIG. 15 depicts chemical structures of 2,5-furandicarboxylic acid (H₂(2,5-FDC)) and 2,4-thiophenedicarboxylic acid (H₂(2,4-TDC)).

[0035] FIG. 16 depicts a comparison of the powder X-ray diffraction (PXRD) analyses conducted on Al(OH)(2,5-FDC), MTV-MOF Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5}, and MOF-323.

[0036] FIG. 17 depicts a ¹H-nuclear magnetic resonance (NMR) analysis conducted on the linkers H₂(2,5-FDC) (top) and H₂(2,4-TDC) (bottom), as well as on the digested (*i.e.*, treated as described in Example 4 below) MTV-MOF Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5} (center). Numbers indicate the relative signal intensities of the respective peaks determined through integration.

[0037] FIG. 18 depicts a representative scanning electron microscopy (SEM) micrograph of the MTV-MOF $\text{Al}(\text{OH})(2,5\text{-FDC})_{0.5}(2,4\text{-TDC})_{0.5}$.

[0038] FIGS. 19A-E depict various representative scanning electron microscopy (SEM) micrographs and energy dispersive X-ray spectroscopy (EDS) analyses of the MTV-MOF $\text{Al}(\text{OH})(2,5\text{-FDC})_{0.5}(2,4\text{-TDC})_{0.5}$.

[0039] FIG. 20 depicts a water sorption analyses on $\text{Al}(\text{OH})(2,5\text{-FDC})$, the MTV-MOF $\text{Al}(\text{OH})(2,5\text{-FDC})_{0.5}(2,4\text{-TDC})_{0.5}$, and MOF-323 at 25 °C (P: partial vapor pressure, P_{sat} : saturation vapor pressure at 25 °C).

[0040] FIG. 21 depicts a low pressure region of the water sorption analyses on $\text{Al}(\text{OH})(2,5\text{-FDC})$, the MTV-MOF $\text{Al}(\text{OH})(2,5\text{-FDC})_{0.5}(2,4\text{-TDC})_{0.5}$, and MOF-323 at 25 °C (P: partial vapor pressure, P_{sat} : saturation vapor pressure at 25 °C).

[0041] FIG. 22 depicts chemical structures of Isophthalic acid (H_2IPA) and 3,5-pyridinedicarboxylic acid ($\text{H}_2(3,5\text{-PynDC})$).

[0042] FIG. 23 depicts a comparison of the powder X-ray diffraction (PXRD) analyses conducted on $\text{Al}(\text{OH})(3,5\text{-PynDC})$, the MTV-MOF $\text{Al}(\text{OH})(3,5\text{-PynDC})_{0.56}(\text{IPA})_{0.44}$, and $\text{Al}(\text{OH})\text{IPA}$.

[0043] FIG. 24 depicts a ^1H -Nuclear magnetic resonance (NMR) analysis conducted on the linkers $\text{H}_2(3,5\text{-PynDC})$ (top) and H_2IPA (bottom), as well as on the digested (*i.e.*, treated as described in Example 5 below) MTV-MOF $\text{Al}(\text{OH})(3,5\text{-PynDC})_{0.56}(\text{IPA})_{0.44}$. Numbers indicate the relative signal intensities of the respective peaks determined through integration.

[0044] FIG. 25 depicts a representative scanning electron microscopy (SEM) micrograph of the MTV-MOF $\text{Al}(\text{OH})(3,5\text{-PynDC})_{0.56}(\text{IPA})_{0.44}$.

[0045] FIGS. 26A-E depict representative scanning electron microscopy (SEM) micrographs and energy dispersive X-ray spectroscopy (EDS) analyses of the MTV-MOF $\text{Al}(\text{OH})(3,5\text{-PynDC})_{0.56}(\text{IPA})_{0.44}$.

[0046] FIG. 27 depicts a water sorption analyses on $\text{Al}(\text{OH})(3,5\text{-PynDC})$, the MTV-MOF $\text{Al}(\text{OH})(3,5\text{-PynDC})_{0.56}(\text{IPA})_{0.44}$, and $\text{Al}(\text{OH})\text{IPA}$ at 25 °C (P: partial vapor pressure, P_{sat} : saturation vapor pressure at 25 °C).

[0047] FIG. 28 depicts a low pressure region of the water sorption analyses on Al(OH)(3,5-PynDC), the MTV-MOF Al(OH)(3,5-PynDC)_{0.56}(2,4-TDC)_{0.44}, and Al(OH)IPA at 25 °C (P: partial vapor pressure, P_{sat}: saturation vapor pressure at 25 °C).

DETAILED DESCRIPTION

[0048] The following description sets forth exemplary methods, parameters and the like. It should be recognized, however, that such description is not intended as a limitation on the scope of the present disclosure but is instead provided as a description of exemplary embodiments.

[0049] Unless contraindicated or noted otherwise, in these descriptions and throughout this specification, the terms “a” and “an” mean one or more, the term “or” means and/or. It is understood that the examples and embodiments described herein are for illustrative purposes only and that various modifications or changes in light thereof will be suggested to persons skilled in the art and are to be included within the spirit and purview of this application and scope of the appended claims.

[0050] In an aspect, provided herein are organic-inorganic hybrid materials, in particular metal-organic frameworks (MOFs), constructed from one or more organic building blocks, particularly for water harvesting applications. MOFs are shown to exhibit desirable properties for water harvesting applications under arid conditions. For example, MOFs may be used to adsorb large amounts of atmospheric water at low relative humidity and successive energy-efficient desorption without loss of porosity. This invention identifies reported as well as new MOFs which are particularly suitable for water harvesting under arid conditions. Furthermore, the invention provides technology to tune and improve the water harvesting properties by employing multivariate (MTV) MOFs. Such metal-organic frameworks may be constructed from a combination of two or more organic units/ligands.

[0051] The described approach for tuning of water sorption properties can be applied in MOF-based water harvesting devices. The harvested water can be used for human consumption or irrigation of crops.

[0052] Additionally, the tuning of water sorption properties by employment of the MOFs described herein can be used in other water sorption based applications, such as in heat pumps, dehumidifiers, adsorption refrigerators, solar cooling systems, dryers, organic light emitting devices and secondary battery devices.

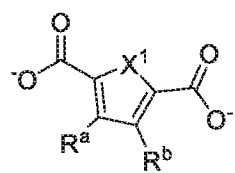
[0053] The following description describes various types of MOFs, including single-linker MOFs and MTV-MOFs, for water sorption applications, in particular water harvesting.

[0054] In some aspects, provided is a MOF comprising repeating cores, wherein the cores comprise secondary building units connected to organic ligands.

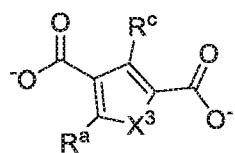
[0055] In some embodiments, the organic ligands comprise one or more linear ditopic moieties, v-shaped ditopic moieties, trigonal tritopic moieties, square or rectangular tetratopic moieties, or tetrahedral tetratopic moieties. In one variation, the organic ligands comprise v-shaped ditopic moieties.

[0056] In some variations of the organic ligands, such moieties comprise 5-membered or 6-membered rings substituted with at least two carboxylate groups. In one variation, such moieties comprise 5-membered or 6-membered rings substituted with two carboxylate groups.

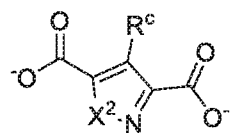
[0057] In certain variations, the organic ligands comprise one or more moieties of Formula I-IX:



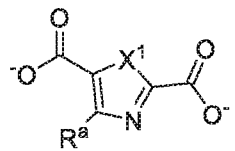
(I), wherein X¹ is NH, O or S, and each of Rᵃ and Rᵇ is independently H or alkyl;



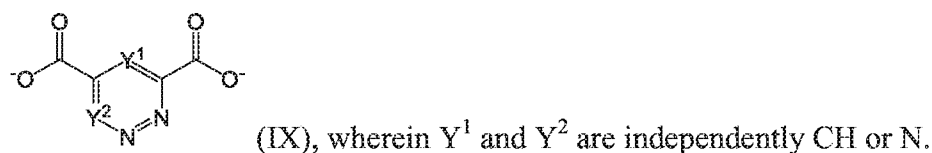
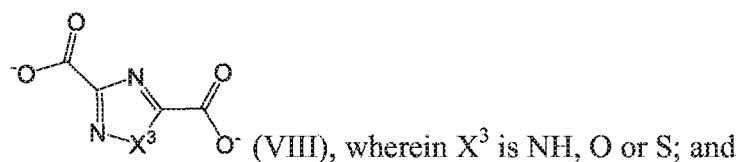
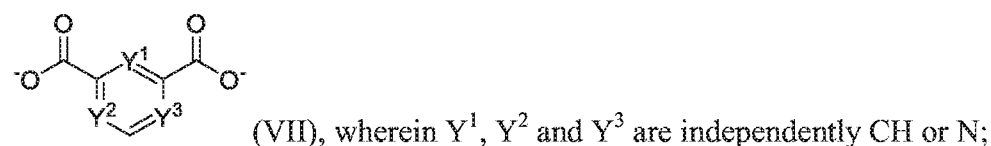
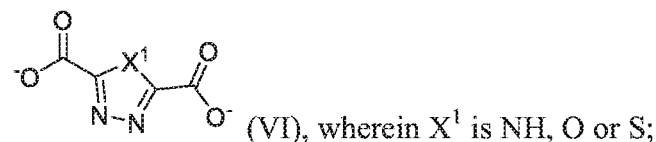
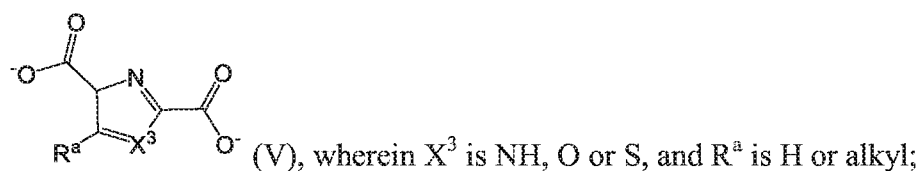
(II), wherein X³ is NH, O or S, and each of Rᵃ and Rᶜ is independently H or alkyl;



(III), wherein X² is NH, O or S, and Rᶜ is H or alkyl;



(IV), wherein X¹ is NH, O or S, and Rᵃ is H or alkyl;



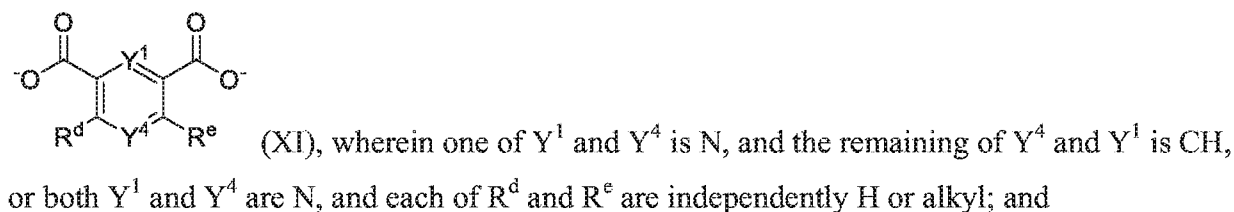
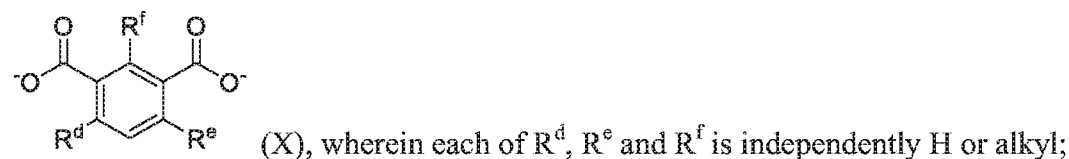
[0058] In one variation of the moiety of Formula (III), X^2 is NH or O, and R^c is H or alkyl.

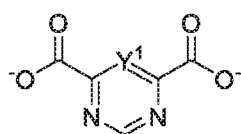
[0059] In one variation of the moiety of Formula (IV), X^1 is NH or O, and R^a is H or alkyl.

[0060] In one variation of the moiety of Formula (V), X^1 is NH or O, and R^a is H or alkyl.

[0061] In one variation of the moiety of Formula (VI), X^1 is NH or O.

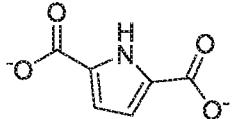
[0062] In certain variations, the organic ligands comprise one or more moieties of Formula (X), (XI), or (XII):

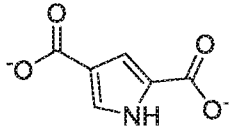


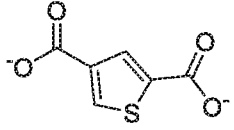


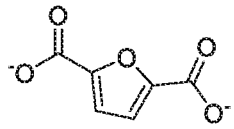
(XII), wherein Y^1 is CH or N.

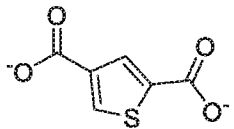
[0063] In some variations of the foregoing, R^a , R^b , R^c , R^d , R^e and R^f are each H.

[0064] In some variations, each organic ligand comprises . In another variation of the foregoing, each secondary building unit comprises aluminum hydroxide. For example, in one variation, the MOF is MOF-313.

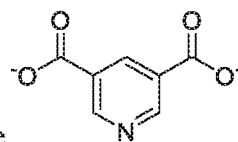
[0065] In some variations, each organic ligand comprises . In another variation of the foregoing, each secondary building unit comprises aluminum hydroxide. For example, in one variation, the MOF is MOF-314.

[0066] In some variations, each organic ligand comprises . In another variation of the foregoing, each secondary building unit comprises aluminum hydroxide. For example, in one variation, the MOF is MOF-323.

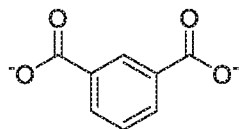
[0067] In some variations, the organic ligands comprise  and

. In another variation of the foregoing, each secondary building unit comprises aluminum hydroxide. For example, in one variation, the MTV-MOF is $Al(OH)(2,5-FDC)(2,4-TDC)$. In certain variations, the MTV-MOF is $Al(OH)(2,5-FDC)_w(2,4-TDC)_v$, wherein $w + v = 1$. In a certain variation, the MTV-MOF is $Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5}$.

[0068] In some variations, the organic ligands comprise



and .



. In another variation of the foregoing, each secondary building unit comprises aluminum hydroxide. For example, in one variation, the MTV-MOF is $\text{Al}(\text{OH})(3,5\text{-PynDC})(\text{IPA})$. In certain variations, the MTV-MOF is $\text{Al}(\text{OH})(3,5\text{-PynDC})_m(\text{IPA})_n$, wherein $m + n = 1$. In a certain variation, the MTV-MOF is $\text{Al}(\text{OH})(3,5\text{-PynDC})_{0.56}(\text{IPA})_{0.44}$.

[0069] FIGS. 1A and 1B provides a non-comprehensive list of compounds that may be used to form organic ligands connected to secondary building units in the MOFs described herein. These linkers can also be combined in systems for water harvesting to expand the realm of MOF materials for water harvesting dramatically.

[0070] In other embodiments, the secondary building units comprise one or more metals or metal-containing complexes. In some variations, the secondary building units form 1D rod-like chains or distinct multinuclear metal clusters. In one variation, each secondary building unit comprises one metal or metal-containing complex.

[0071] In certain variations, each secondary building unit comprises zirconium, nickel, iron, copper, manganese, aluminum, magnesium, calcium, strontium, barium, titanium, zinc, indium, cadmium, hafnium, lead, cobalt, or chromium, or a complex thereof. In one variation, each secondary building unit comprises aluminum or an aluminum-containing complex.

[0072] As a result of the cooperative nature of water condensation in MOF,⁹ the use of the MOFs described herein, including the MTV-MOFs, could enable shifting of the P/P_0 value at which a steep step is observed in the water vapor sorption isotherm. Further, the total water uptake of a MOF with voluminous ligands can be improved by “doping” it with less bulky ligands due to the resulting increase in the pore volume of the MTV-system. Also, linkers otherwise not forming a MOF can be included in a MTV-MOF, thus expanding the realm of water harvesting materials enormously.

[0073] This approach can be employed for fine-tuning of a system in such a way that optimal properties for a desired water sorption application can be achieved. Thus, the described invention enables the synthesis of water harvesting materials on demand.

[0074] In some variations, provided herein are Al-MOFs exhibiting 1D rod-like SBUs constructed from V-shaped molecules. Such MTV-approach can also be employed for other MOF-based water harvesting systems constructed from other metal cations and linkers with other geometries.

[0075] In certain variations, the secondary building units are connected to the organic ligands through the oxygen atoms of the carboxylate groups in the organic ligands via a cis-edge-shared octahedra geometry or a trans-edge-shared octahedra geometry.

[0076] In other aspects, provided are methods of producing the MOFs described herein. In some embodiments, the method comprises combining one or more of the compounds set forth in FIGS. **1A** and **1B** with a metal solution under basic conditions, at elevated temperatures. In some variations, the metal solution comprises any of the metals described herein for the MOFs. In other variations, the compounds in FIGS. **1A** and **1B** and the metal solution are combined with a basic solution, such as sodium hydroxide. A solid is obtained from the reaction mixture, and isolated. The isolated product can then be characterized using any suitable methods or techniques known in the art, including, for example, by x-ray diffraction (*e.g.*, powder x-ray diffraction).

[0077] In yet other aspects, provided are methods and systems for water harvesting using one or more of the MOFs described herein. In one aspect, provided is a method of water harvesting, comprising: adsorbing water from ambient atmosphere using a water-harvesting system, wherein the water-harvesting system comprises one or more of the MOFs described herein; desorbing vapor from the one or more MOFs; and collecting water from the vapor. In another aspect, provided is a water-harvesting system, comprising an adsorbent layer comprising one or more of the MOFs described herein. In some variations of the foregoing methods and systems for water harvesting, the MOFs used are MTV-MOFs.

[0078] In some variations of the foregoing, the water-harvesting system is a passive device, in which sunlight-driven desorption of water leads to saturation in a closed environment which, consequently, leads to water condensation. In other variations, the water-harvesting system is an active device, in which a condenser is needed to collect the water. The condenser may be, for example, adjacent to the adsorbent layer in the water-harvesting system.

ENUMERATED EMBODIMENTS

[0079] The following enumerated embodiments are representative of some aspects of the invention.

1. A composition comprising a multivariate metal-organic framework (MTV-MOF) of inorganic metal clusters (secondary building units, SBUs) and a combination of two or more different organic units (linkers).
2. The composition of embodiment 1, wherein the linkers are selected from v-shaped, trigonal or tritopic linkers; or square/rectangular-shaped tetratopic linkers; or tetrahedral tetratopic linkers.
3. The composition of embodiment 1, wherein at least one of the linkers is a v-shaped linker disclosed herein.
4. The composition of embodiment 1, wherein the metal clusters are infinite 1D rod-like chains or distinct multinuclear metal clusters.
5. The composition of any one of embodiments 1 to 4, wherein the MOF metal is selected from zirconium, nickel, iron, copper, manganese and aluminum, and/or from magnesium, calcium, strontium, barium, titanium, zinc, indium, cadmium, hafnium, lead, cobalt, and chromium.
6. A device such as a water harvester, comprising a moisture sorption unit comprising the composition of any one of embodiments 1 to 5.
7. A method of using the composition of any one of embodiments 1 to 6, comprising containing in, storing in and/or extracting from the composition a predetermined gas or fluid, such as CO₂, H₂O, H₂, CH₄, C₂H₄, C₂H₂, etc.
8. A method of using the composition of any one of embodiments 1 to 6, comprising utilizing the composition for water harvesting or water purification applications.
9. A water harvester comprising a moisture sorption unit comprising a MOF of formula Al(OH)(2,5-PyIDC), wherein 2,5-PyIDC is 2,5-pyrroledicarboxylate.
10. A method of using the harvester of embodiment 9, for water harvesting or water purification applications.
11. A metal-organic framework (MOF), comprising repeating cores, wherein the cores comprise secondary building units connected to organic ligands,

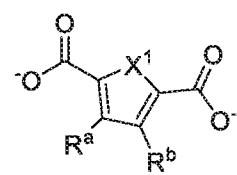
wherein the secondary building units comprise one or more metals or metal-containing complexes,

wherein the organic ligands comprise one or more linear ditopic moieties, v-shaped ditopic moieties, trigonal tritopic moieties, square or rectangular tetratopic moieties, or tetrahedral tetratopic moieties, wherein the moieties comprise 5-membered or 6-membered rings substituted with at least two carboxylate groups, and

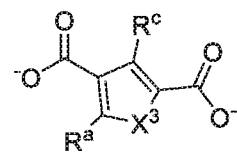
wherein the secondary building units are connected to the organic ligands through the oxygen atoms of the carboxylate groups in the organic ligands.

12. The MOF of embodiment 11, wherein the secondary building units form 1D rod-like chains or distinct multinuclear metal clusters.

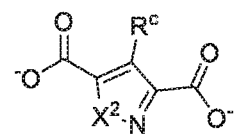
13. The MOF of embodiment 11 or 12, wherein the organic ligands comprise one or more moieties of Formula (I)-(XII):



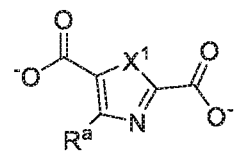
(I), wherein X¹ is NH, O or S, and each of R^a and R^b is independently H or alkyl;



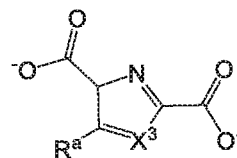
(II), wherein X³ is NH, O or S, and each of R^a and R^c is independently H or alkyl;



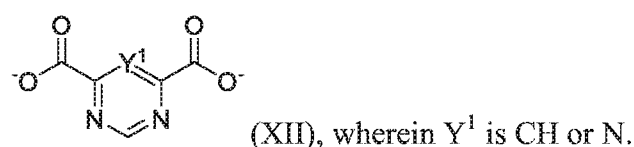
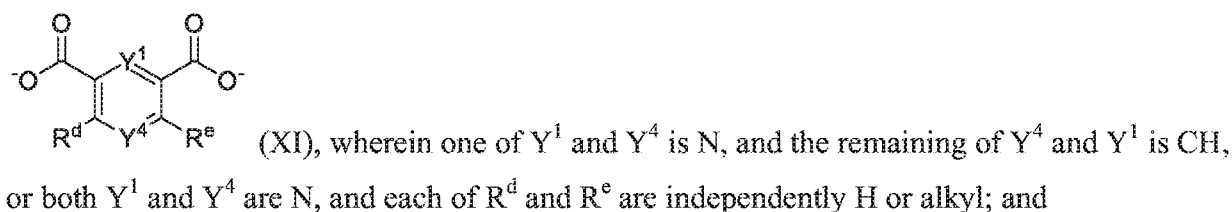
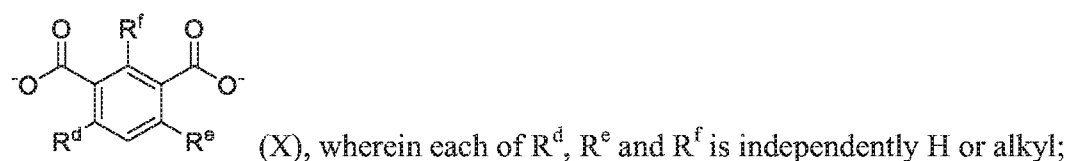
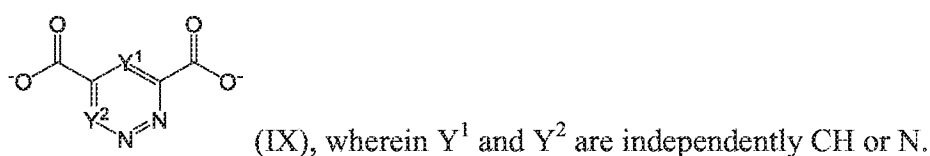
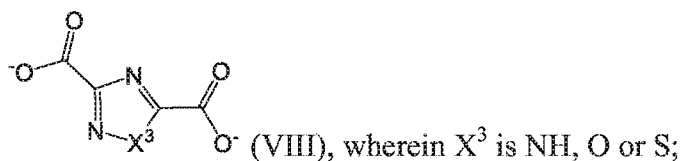
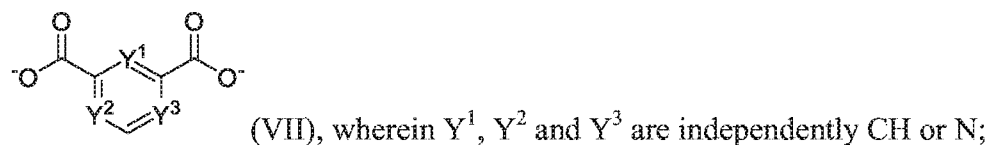
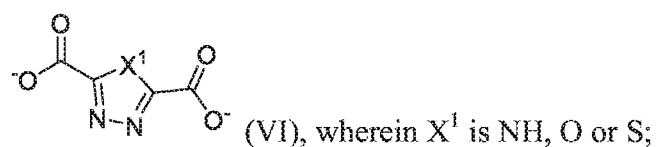
(III), wherein X² is NH, O or S, and R^c is H or alkyl;



(IV), wherein X¹ is NH, O or S, and R^a is H or alkyl;



(V), wherein X³ is NH, O or S, and R^a is H or alkyl;



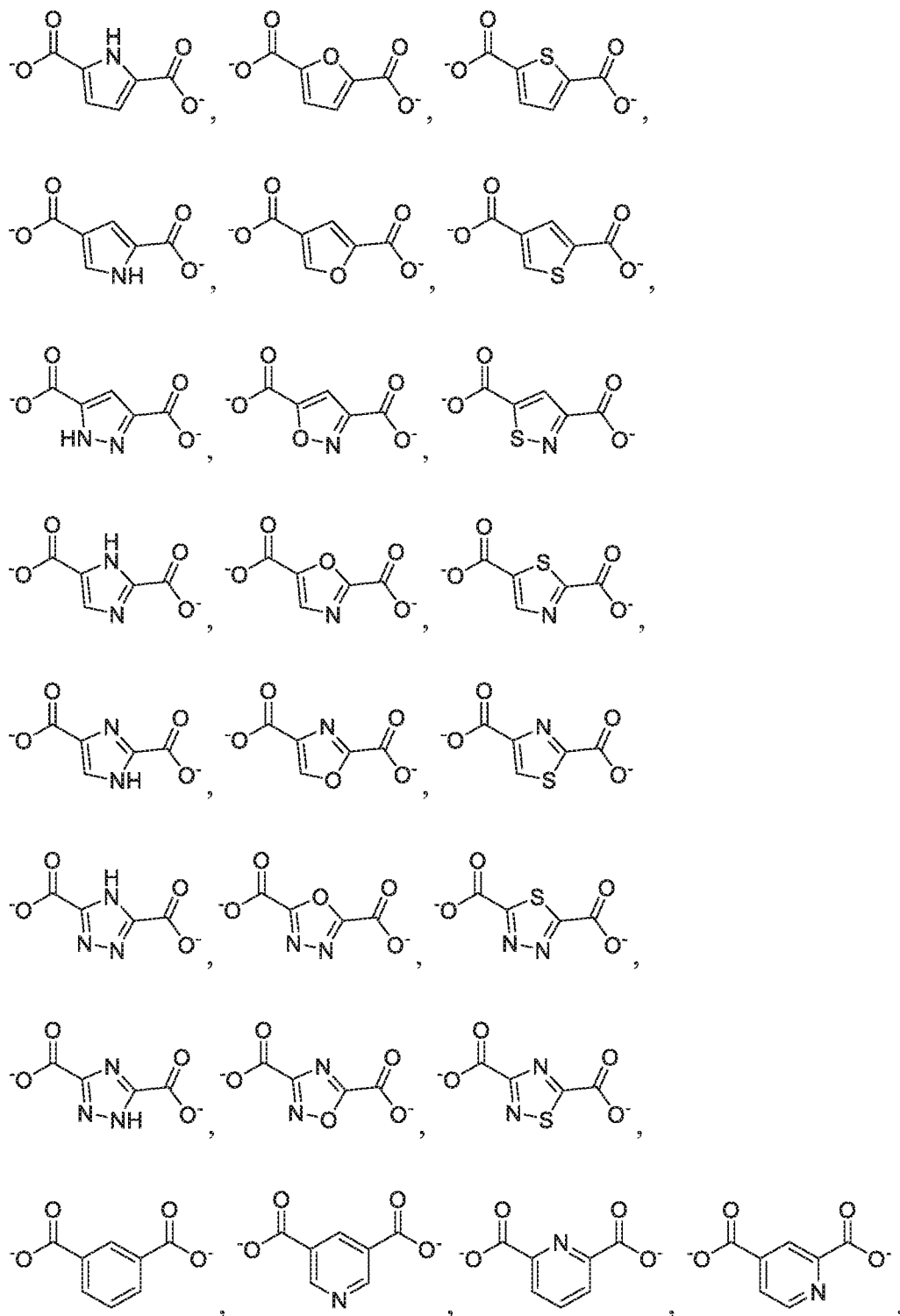
14. The MOF of embodiment 13, wherein the organic ligands comprise one or more moieties of Formula (I)-(VII).

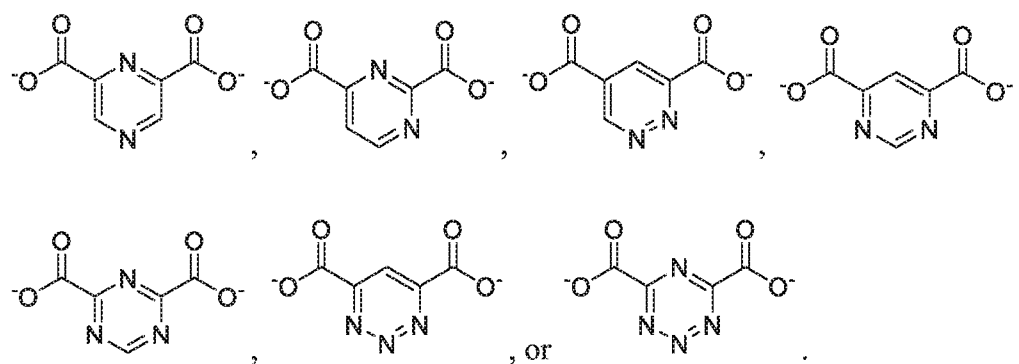
15. The MOF of embodiment 13, wherein the organic ligands comprise one or more moieties of Formula (VIII) or (IX).

16. The MOF of embodiment 13, wherein the organic ligands comprise one or more moieties of Formula (X), (XI) or (XII).

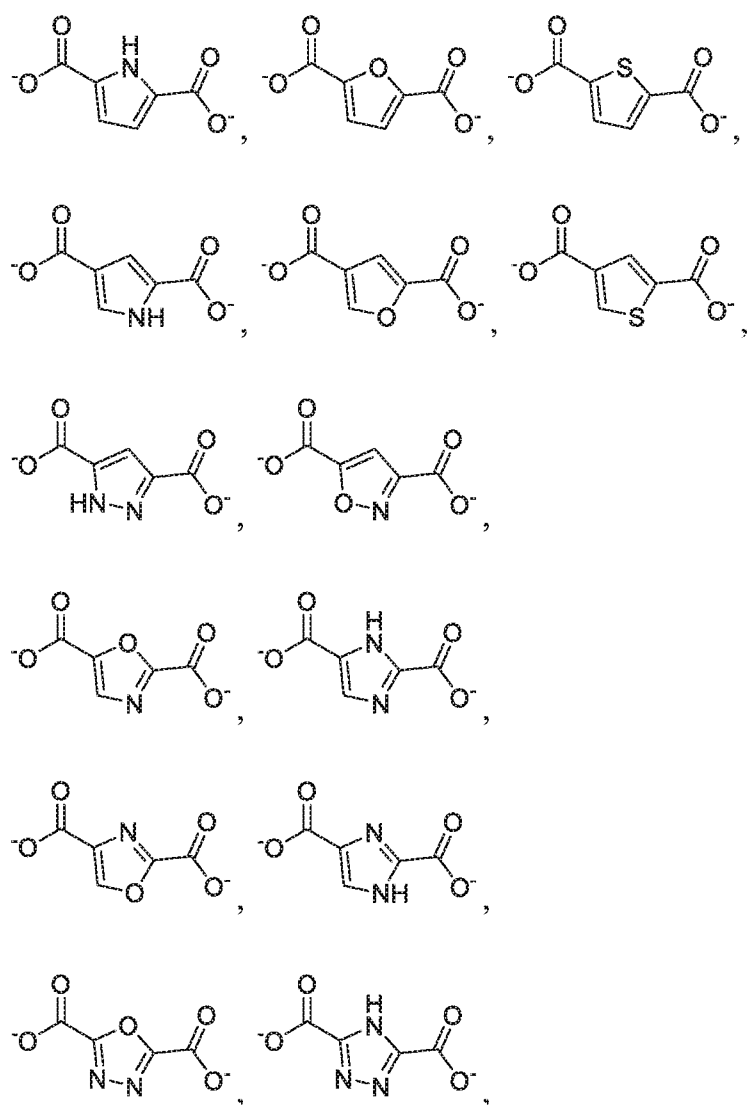
17. The MOF of any one of embodiments 13 to 16, wherein R^a , R^b , R^c , R^d , R^e and R^f are each H.

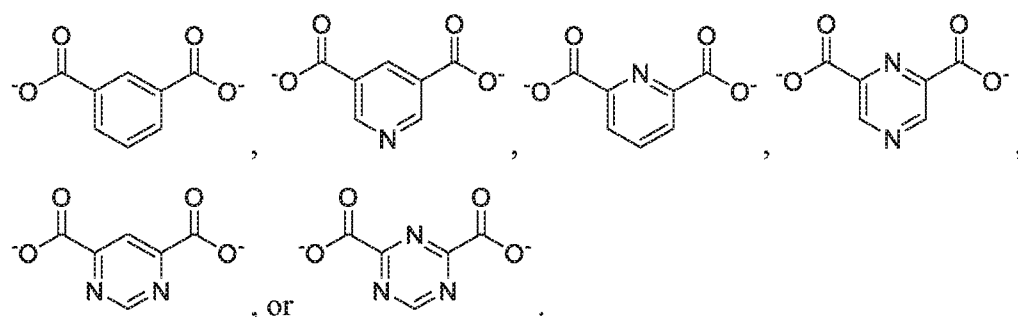
18. The MOF of embodiment 11 or 12, wherein the organic ligands comprise one or more moieties:



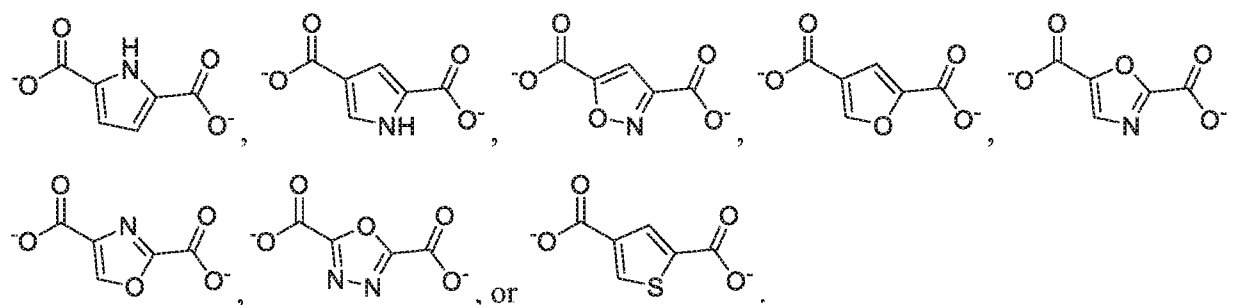


19. The MOF of embodiment 11 or 12, wherein the organic ligands comprise one or more moieties:





20. The MOF of embodiment 11 or 12, wherein the organic ligands comprise one or more moieties:



21. The MOFs of any one of embodiments 18 to 20, wherein the organic ligands comprise one moiety.

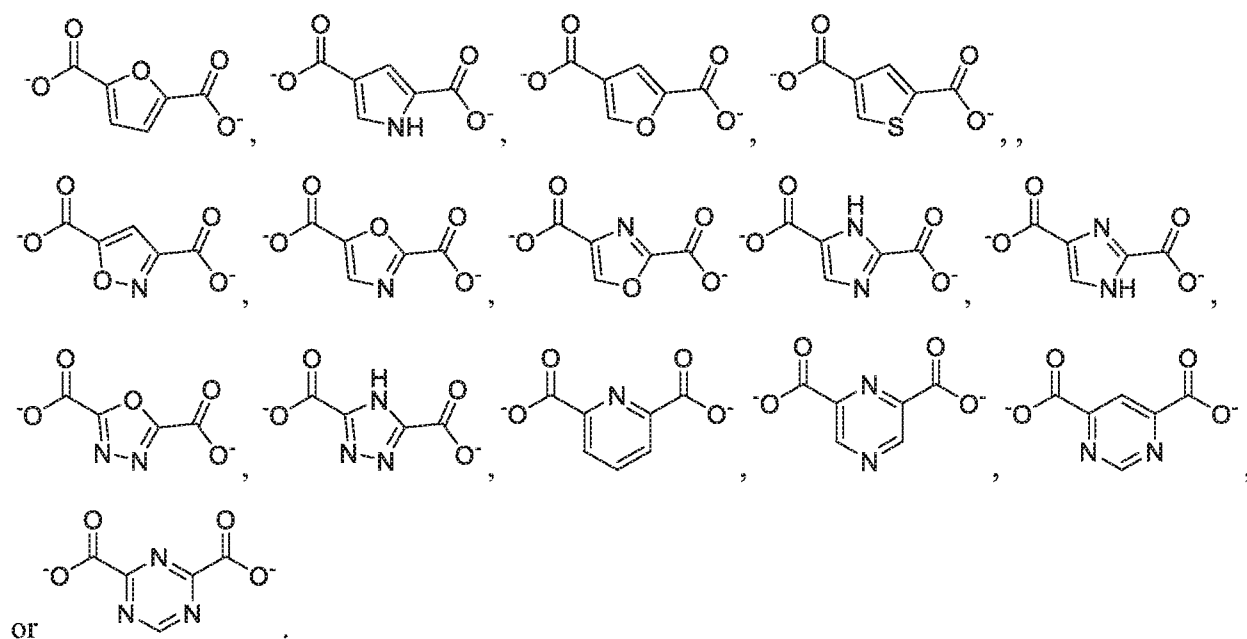
22. The MOFs of any one of embodiments 18 to 20, wherein the organic ligands comprise two moieties.

23. The MOFs of any one of embodiments 18 to 20, wherein the organic ligands comprise a plurality of moieties.

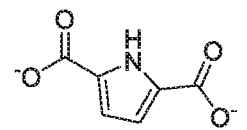
24. The MOF of any one of embodiments 11 to 17, wherein each organic ligand comprises one moiety of Formula (I)-(XII), and the MOF is a single-linker metal-organic framework.

25. The MOF of embodiment 24, wherein each organic ligand comprises one moiety of Formula (I)-(VII), and the MOF is a single-linker metal-organic framework.

26. The MOF of embodiment 24 wherein each organic ligand comprises:



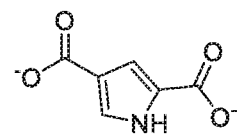
27. The MOF of embodiment 24, wherein each organic ligand comprises



28. The MOF of embodiment 27, wherein each secondary building unit comprises aluminum hydroxide.

29. The MOF of embodiment 28, wherein the MOF is MOF-313.

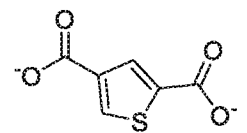
30. The MOF of embodiment 24, wherein each organic ligand comprises



31. The MOF of embodiment 30, wherein each secondary building unit comprises aluminum hydroxide.

32. The MOF of embodiment 31, wherein the MOF is MOF-314.

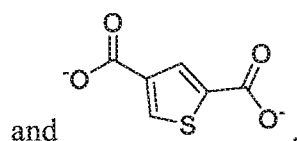
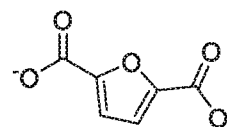
33. The MOF of embodiment 24, wherein each organic ligand comprises



34. The MOF of embodiment 33, wherein each secondary building unit comprises aluminum hydroxide.

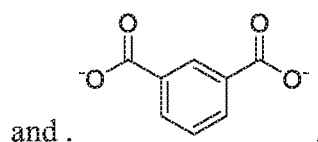
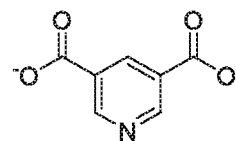
35. The MOF of embodiment 34, wherein the MOF is MOF-323.
36. The MOF of any one of embodiments 11 to 20, wherein the MOF is a multivariate metal-organic framework (MTV-MOF).
37. The MOF of embodiment 36, wherein the organic ligands comprises a plurality of moieties of Formula (I)-(VII).
38. The MOF of embodiment 36, wherein the organic ligands comprises two moieties of Formula (I)-(VII).
39. The MOF of embodiment 36, wherein the organic ligands comprises a plurality of moieties of Formula (I)-(XII).
40. The MOF of embodiment 36, wherein the organic ligands comprises two moieties of Formula (I)-(XII).

41. The MOF of embodiment 36, wherein the organic ligands comprise



42. The MOF of embodiment 41, wherein each secondary building unit comprises aluminum hydroxide.
43. The MOF of embodiment 42, wherein the MOF is $\text{Al}(\text{OH})(2,5\text{-FDC})(2,4\text{-TDC})$; or $\text{Al}(\text{OH})(2,5\text{-FDC})_w(2,4\text{-TDC})_v$, wherein $w + v = 1$; or $\text{Al}(\text{OH})(2,5\text{-FDC})_{0.5}(2,4\text{-TDC})_{0.5}$.

44. The MOF of embodiment 36, wherein the organic ligands comprise



45. The MOF of embodiment 44, wherein each secondary building unit comprises aluminum hydroxide.

46. The MOF of embodiment 45, wherein the MOF is $\text{Al}(\text{OH})(3,5\text{-PynDC})(\text{IPA})$, or optionally $\text{Al}(\text{OH})(3,5\text{-PynDC})_m(\text{IPA})_n$, wherein $m + n = 1$; or $\text{Al}(\text{OH})(3,5\text{-PynDC})_{0.56}(\text{IPA})_{0.44}$.
47. The MOF of any one of embodiments 1 to 46, wherein each secondary building unit comprises zirconium, nickel, iron, copper, manganese, aluminum, magnesium, calcium, strontium, barium, titanium, zinc, indium, cadmium, hafnium, lead, cobalt, or chromium, or a complex thereof.
48. The MOF of any one of embodiments 11 to 47, wherein each secondary building unit comprises aluminum or an aluminum-containing complex.
49. The MOF of any one of embodiments 11 to 48, wherein the secondary building units form 1D rod-like chains, and the organic ligands comprise one or more v-shaped ditopic moieties.
50. The MOF of embodiment 49, wherein the secondary building units are connected to the organic ligands through the oxygen atoms of the carboxylate groups in the organic ligands via a cis-edge-shared octahedra geometry or a trans-edge-shared octahedra geometry.
51. A method of water harvesting, comprising:
adsorbing water from ambient atmosphere using a water-harvesting system, wherein the water-harvesting system comprises one or more MOFs of any one of embodiments 11 to 50;
desorbing vapor from the one or more MOFs; and
collecting water from the vapor.
52. The method of embodiment 51, wherein the water is collected from the vapor using a condenser.
53. A water harvesting system, comprising an adsorbent layer comprising one or more MOFs of any one of embodiments 11 to 50.
54. The system of embodiment 53, further comprising a condenser adjacent to the adsorbent layer.
55. The system of embodiment 53, wherein the system is a passive water harvesting system.
56. A MOF of any one of embodiments 11 to 50 for use in water harvesting or water purification.

57. A MOF of any one of embodiments 11 to 50 for use in heat pumps, dehumidifiers, adsorption refrigerators, solar cooling systems, dryers, organic light emitting devices, or secondary battery devices.

58. A method of using the MOF of any one of embodiments 11 to 50, comprising containing in, storing in and/or extracting from the MOF a predetermined gas or fluid.

59. The method of embodiment 58, wherein the gas or fluid is CO₂, H₂O, H₂, CH₄, C₂H₄, or C₂H₂, or any combinations thereof.

EXAMPLES

[0080] The following Examples are merely illustrative and are not meant to limit any aspects of the present disclosure in any way.

Example 1

Synthesis and Characterization of MOF-313

[0081] To illustrate an exemplary way to synthesize 1D rod-like SBU based MOFs, the synthesis and characterization of Al(OH)(2,5-PyIDC) (2,5-PyIDC = 2,5-pyrroledicarboxylate), a MOF suitable for water sorption applications, is described here. This compound is also referred to as MOF-313.

Synthesis:

[0082] 77.5 mg of 2,5-pyrroledicarboxylic acid (0.5 mmol) was dissolved in 4.5 mL aqueous NaOH solution (0.139 M) under heating. Then, 500 μ L aqueous AlCl₃ solution (1 M) was added. The resulting precipitation was dissolved under heating and the solution was placed in an oven preheated to 100 °C. After two hours, the precipitation was filtrated off. The solid was washed five times with water for one day, and five times with methanol for another day. Removal of the solvent under reduced pressure resulted in 15 mg Al(OH)(2,5-PyIDC).

PXRD Characterization:

[0083] The structure (FIGS. 2 and 3) of the synthesized MOF was verified based on powder x-ray diffraction (PXRD) analysis (FIG. 4). The structure was shown to be constructed from 1D rod-like SBUs connected by 2,5-pyrroledicarboxylate linker ligands.

Porosity and Surface Area:

[0084] In order to analyze the porosity and surface area of the metal-organic framework, a nitrogen isotherm at 77 K was collected (FIG. 5). Based on the BET model, the surface area was estimated to be 1051 m²/g.

Water Sorption Properties:

[0085] The water sorption isotherm exhibited a steep step at $P/P_0 = 0.12$ (FIG. 6). Water vapor uptake started to level off at $P/P_0 = 0.15$ and reached approximately 480 cm³/g (STP) at full saturation of the material. Importantly, only minimal or in the optimal case no hysteresis was observed.

Example 2

Synthesis and Characterization of MOF-314

[0086] To illustrate an exemplary way to synthesize 1D rod-like SBU based MOFs, the synthesis and characterization of Al(OH)(2,4-PyIDC) (2,4-PyIDC = 2,4-pyrroledicarboxylate), a MOF suitable for water sorption applications, is described here. This compound is also referred to as MOF-314.

Synthesis:

[0087] 2,4-pyrroledicarboxylic acid [H₂(2,4-PyIDC); 75.0 mg] was completely dissolved in aqueous NaOH solution (0.43 M, 14.517 mL). Then, aqueous AlCl₃ solution (1M, 0.483 mL) was added which resulted in white precipitation. The resulting suspension was heated to 100 °C in a pre-heated oven and the solution became clear after a few minutes. The reaction was stopped after 6 hours. The resulting off-white solid was filtered and washed three times with H₂O and three times with methanol over a period of one day each. After heating at 120 °C under vacuum ($\sim 10^{-3}$ mbar) for 6 hours, activated MOF-314 [Al(OH)(2,4-PyIDC); 35 mg] was obtained.

PXRD Characterization:

[0088] The structure (FIG. 7) of the synthesized MOF was verified based on powder x-ray diffraction (PXRD) analysis (FIG. 8). The structure was shown to be constructed from 1D rod-like SBUs connected by 2,4-pyrroledicarboxylate linker ligands.

Porosity and Surface Area:

[0089] In order to analyze the porosity and surface area of the metal-organic framework, a nitrogen isotherm at 77 K was collected (FIG. 9). Based on the BET model, the surface area was estimated to be 912 m²/g.

Water Sorption Properties:

[0090] The water sorption isotherm exhibited a steep step at $P/P_0 = 0.1$ (FIG. 10). Water vapor uptake started to level off at $P/P_0 = 0.2$ and reached approximately 535 cm³/g (STP) at full saturation of the material. Importantly, only minimal or in the optimal case no hysteresis was observed.

Example 3**Synthesis and Characterization of MOF-323**

[0091] To illustrate an exemplary way to synthesize 1D rod-like SBU based MOFs, the synthesis and characterization of Al(OH)(2,4-TDC) (2,4-TDC = 2,4-thiophenedicarboxylate), a MOF suitable for water sorption applications, is described here. This compound is also referred to as MOF-323.

Synthesis:

[0092] 2,4-thiophenedicarboxylic acid (H₂(2,4-TDC); 98.9 mg) was dissolved in KOH solution (0.139 M, 6.187 mL) and ethanol (3.239 mL). Then, aqueous AlCl₃ solution (1 M, 0.574 mL) was added. The resulting clear solution was heated to 100 °C in a pre-heated oven. After 24 hours, the resulting white precipitation was filtered and washed three times with H₂O and three times with methanol over a period of one day each. After heating at 60 °C under vacuum ($\sim 10^{-3}$ mbar) for 6 hours, activated MOF-323 (Al(OH)(2,4-TDC); 60 mg) was obtained.

PXRD Characterization:

[0093] The structure (FIG. 11) of the synthesized MOF was verified based on powder x-ray diffraction (PXRD) analysis (FIG. 12). The structure was shown to be constructed from 1D rod-like SBUs connected by 2,4-thiophenedicarboxylate linker ligands.

Porosity and Surface Area:

[0094] In order to analyze the porosity and surface area of the metal-organic framework, a nitrogen isotherm at 77 K was collected (FIG. 13). Based on the BET model, the surface area was estimated to be 1101 m²/g.

Water Sorption Properties:

[0095] The water sorption isotherm exhibited a steep step at $P/P_0 = 0.16$ (FIG. 14). Water vapor uptake started to level off at $P/P_0 = 0.21$ and reached approximately 470 cm³/g (STP) at full saturation of the material. Importantly, only minimal or in the optimal case no hysteresis was observed.

Example 4

Synthesis and Characterization of MTV-MOF

[0096] To illustrate an exemplary way to synthesize an exemplary MTV-MOF system, the synthesis and characterization of Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5} (2,5-FDC = 2,5-furandicarboxylate; 2,4-TDC = 2,4-thiophenedicarboxylate; FIG. 15), a MTV-MOF suitable for water sorption applications, is described here.

Synthesis of the MTV-MOF:

[0097] 2,5-furandicarboxylic acid (H₂(2,5-FDC); 44.9 mg) and 2,4-thiophenedicarboxylic acid (H₂(2,4-TDC); 49.5 mg) were dissolved in aqueous KOH solution (0.139 M, 6.187 mL) and ethanol (3.239 mL) (FIG. 15). Then, aqueous AlCl₃ solution (1 M, 0.574 mL) was added. The resulting solution was heated to 100 °C in a pre-heated oven. After 24 hours, the resulting white precipitation was filtered and washed three times with H₂O and three times with methanol over a period of one day each. After heating at 120 °C under vacuum (~10⁻³ mbar) for 6 hours, the activated MTV-MOF Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5} (35 mg) was obtained.

Synthesis of the single-linker MOFs used for comparison:

[0098] Al(OH)(2,5-FDC) (also known as MIL-160) was synthesized by dissolving H₂(2,5-FDC) (78.0 mg) in aqueous KOH solution (0.079 M, 9.5 mL) and adding aqueous AlCl₃ solution (1 M, 0.5 mL). The resulting solution was heated at 100 °C overnight. The resulting white precipitation was filtered and washed three times with H₂O and three times with methanol over a period of one day each. After heating at 120 °C under vacuum (~10⁻³ mbar) for 6 hours, activated Al(OH)(2,5-FDC) (40 mg) was obtained. Al(OH)(2,4-TDC) (MOF-323) was synthesized as described above.

PXRD characterization:

[0099] The synthesized MTV-MOF was verified based on powder x-ray diffraction (PXRD) analysis, and compared with Al(OH)(2,5-FDC) and MOF-323 (FIG. 16).

NMR analysis:

[0100] Washed and activated samples were subjected to NaOD solution (5% in D₂O) prior to nuclear magnetic resonance (NMR) analysis. Under these conditions, the aluminum MTV-MOF dissolved into its components, and solution-based ¹H-NMR analysis was employed to analyze its linker molecule composition (FIG. 17).

Scanning electron microscopy (SEM) micrograph:

[0101] The overall crystal shape and morphology in an SEM micrograph of the MTV-MOF Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5} (FIG. 18) indicated phase-purity.

Scanning electron microscopy (SEM) micrograph and energy dispersive X-ray spectroscopy (EDS) analysis:

[0102] An SEM micrograph and EDS analysis of the MTV-MOF Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5} (FIGS. 19A-E) indicated even distribution of the elements carbon (C), oxygen (O), aluminum (Al), and sulfur (S) in all crystals. Together with the NMR data, this indicated that both linkers are present and evenly distributed in all crystals.

Water sorption properties:

[0103] Water sorption analyses on Al(OH)(2,5-FDC), the MTV-MOF Al(OH)(2,5-FDC)_{0.5}(2,4-TDC)_{0.5}, and MOF-323 at 25 °C (FIGS. 20 and 21) indicated that MTV-systems can be used to tune the water sorption properties of MOFs.

Example 5**Synthesis and Characterization of MTV-MOF**

[0104] To illustrate an exemplary way to synthesize another exemplary MTV-MOF system, the synthesis and characterization of Al(OH)(3,5-PynDC)_{0.56}(IPA)_{0.44} (3,5-PynDC = 3,5-pyridinedicarboxylate; IPA = isophthalate; FIG. 22), a MTV-MOF suitable for water sorption applications, is described here.

Synthesis of the MTV-MOF:

[0105] Isophthalic acid (H_2IPA ; 41.5 mg) and 3,5-pyridinedicarboxylic acid ($H_2(3,5-PynDC)$; 41.8 mg) were dissolved in H_2O (4 mL) and ethanol (5 mL) (FIG. 22). Then, aqueous $AlCl_3$ (1 M, 0.5 mL) and aqueous urea solution (2 M, 0.5 mL) were added. The resulting solution was heated to 100 °C in a pre-heated oven. After 24 hours, the resulting white precipitation was filtered and washed three times with H_2O and three times with methanol over a period of one day each. After heating at 120 °C under vacuum ($\sim 10^{-3}$ mbar) for 6 hours, the activated MTV-MOF $Al(OH)(3,5-PynDC)_{0.56}(IPA)_{0.44}$ (70 mg) was obtained..

Synthesis of the single-linker MOFs used for comparison:

[0106] To synthesize $Al(OH)(3,5-PynDC)$ (also known as CAU-10-pyridine), 3,5-pyridinedicarboxylic acid ($H_2(3,5-PynDC)$; 83.5 mg) was dissolved in H_2O (4 mL) and ethanol (5 mL). Then, aqueous $AlCl_3$ (1 M, 0.5 mL) and aqueous urea solution (2 M, 0.5 mL) was added. The resulting solution was heated to 100 °C in a pre-heated oven. After 24 hours, the resulting white precipitation was filtered and washed three times with H_2O and three times with methanol over a period of one day each. After heating at 120 °C under vacuum ($\sim 10^{-3}$ mbar) for 6 hours, $Al(OH)(3,5-PynDC)$ (70 mg) was obtained. $Al(OH)IPA$ (also known as CAU-10) was synthesized based on a previously reported procedure.¹⁶

PXRD characterization:

[0107] The synthesized MTV-MOF was verified based on powder x-ray diffraction (PXRD) analysis, and compared with $Al(OH)(3,5-PynDC)$ and $Al(OH)IPA$ (FIG. 23).

NMR analysis:

[0108] Washed and activated samples were subjected to NaOD solution (5% in D_2O) prior to nuclear magnetic resonance (NMR) analysis. Under these conditions, the aluminum MTV-MOF dissolved into its components, such that solution-based ^1H -NMR analysis could be employed to analyze its linker molecule composition (FIG. 24).

Scanning electron microscopy (SEM) micrograph and energy dispersive X-ray spectroscopy (EDS) analysis:

[0109] An SEM micrograph and EDS analysis of the MTV-MOF $Al(OH)(3,5-PynDC)_{0.56}(IPA)_{0.44}$ (FIGS. 25 and 26A-E) indicated even distribution of the elements carbon

(C), oxygen (O), aluminum (Al), and sulfur (S) in all crystals. Together with the NMR data, this indicated that both linkers are present and evenly distributed in all crystals.

Water sorption properties:

[0110] Water sorption analyses on Al(OH)(3,5-PynDC), the MTV-MOF Al(OH)(3,5-PynDC)_{0.56}(IPA)_{0.44}, and Al(OH)IPA at 25 °C (FIGS. 27 and 28) indicated that MTV-systems can be used to tune the water sorption properties of MOFs.

[0111] References

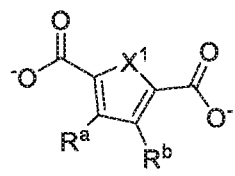
- (1) World Economic Forum. *The Global Risks Report 2018, 13th Edition*; 2018.
- (2) Mekonnen, M. M.; Hoekstra, A. Y. Four Billion People Facing Severe Water Scarcity. *Sci. Adv.* **2016**, 2 (2), e1500323–e1500323.
- (3) Vorosmarty, C. J. Global Water Resources: Vulnerability from Climate Change and Population Growth. *Science* (80-.). **2000**, 289 (5477), 284–288.
- (4) Wada, Y.; Gleeson, T.; Esnault, L. Wedge Approach to Water Stress. *Nat. Geosci.* **2014**, 7 (9), 615–617.
- (5) Wahlgren, R. V. Atmospheric Water Vapour Processor Designs for Potable Water Production: A Review. *Water Res.* **2001**, 35 (1), 1–22.
- (6) Klemm, O.; Schemenauer, R. S.; Lummerich, A.; Cereceda, P.; Marzol, V.; Corell, D.; Van Heerden, J.; Reinhard, D.; Gherezghiher, T.; Olivier, J.; et al. Fog as a Fresh-Water Resource: Overview and Perspectives. *Ambio* **2012**, 41 (3), 221 – 234.
- (7) Tomaszekiewicz, M.; Abou Najm, M.; Beysens, D.; Alameddine, I.; El-Fadel, M. Dew as a Sustainable Non-Conventional Water Resource: A Critical Review. *Environ. Rev.* **2015**, 23 (4), 425 – 442.
- (8) Cells, B. S.; Johansson, M. B.; Johansson, E. Accepted Article. No. Iv.
- (9) Furukawa, H.; Gándara, F.; Zhang, Y.-B.; Jiang, J.; Queen, W. L.; Hudson, M. R.; Yaghi, O. M. Water Adsorption in Porous Metal – Organic Frameworks and Related Materials. *J. Am. Chem. Soc.* **2014**, 136 (11), 4369 – 4381.
- (10) Fathieh, F.; Kalmutzki, M. J.; Kapustin, E. A.; Waller, P. J.; Yang, J.; Yaghi, O. M. Practical Water Production from Desert Air. *Sci. Adv.* **2018**, 4 (6), eaat3198.
- (11) Kalmutzki, M. J.; Diercks, C. S.; Yaghi, O. M. Metal-Organic Frameworks for Water Harvesting from Air. *Adv. Mater.* **2018**, 1704304, 1704304.

- (12) Deng, H.; Doonan, C. J.; Furukawa, H.; Ferreira, R. B.; Towne, J.; Knobler, C. B.; Wang, B.; Yaghi, O. M. Multiple Functional Groups of Varying Ratios in Metal-Organic Frameworks. *Science* (80-.). **2010**, 327 (5967), 846 – 850.
- (13) Reinsch, H.; Van Der Veen, M. A.; Gil, B.; Marszalek, B.; Verbiest, T.; De Vos, D.; Stock, N. Structures, Sorption Characteristics, and Nonlinear Optical Properties of a New Series of Highly Stable Aluminum MOFs. *Chem. Mater.* **2013**, 25 (1), 17 – 26.
- (14) Cadiau, A.; Lee, J. S.; Damasceno Borges, D.; Fabry, P.; Devic, T.; Wharmby, M. T.; Martineau, C.; Foucher, D.; Taulelle, F.; Jun, C. H.; et al. Design of Hydrophilic Metal Organic Framework Water Adsorbents for Heat Reallocation. *Adv. Mater.* **2015**, 27 (32), 4775 – 4780.
- (15) Tschense, C. B. L.; Reimer, N.; Hsu, C. W.; Reinsch, H.; Siegel, R.; Chen, W. J.; Lin, C. H.; Cadiau, A.; Serre, C.; Senker, J.; et al. New Group 13 MIL-53 Derivates Based on 2,5-Thiophenedicarboxylic Acid. *Zeitschrift für Anorg. und Allg. Chemie* **2017**, 643 (21), 1600 – 1608.
- (16) Lenzen, D.; Bendix, P.; Reinsch, H.; Fröhlich, D.; Kummer, H.; Möllers, M.; Hügenell, P. C.; Gläser, R.; Henninger, S.; Stock, N.; Scalable Green Synthesis and Full-Scale Test of the Metal-Organic Framework CAU-10-H for Use in Adsorption-Driven Chillers. *Adv. Mater.* **30**, 1705869 (2018).

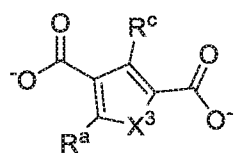
CLAIMS

What is claimed is:

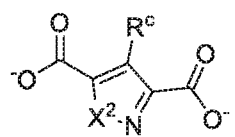
1. A metal-organic framework (MOF), comprising repeating cores, wherein the cores comprise secondary building units connected to organic ligands,
wherein the secondary building units comprise one or more metals or metal-containing complexes,
wherein the organic ligands comprise one or more linear ditopic moieties, v-shaped ditopic moieties, trigonal tritopic moieties, square or rectangular tetratopic moieties, or tetrahedral tetratopic moieties, wherein the moieties comprise 5-membered or 6-membered rings substituted with at least two carboxylate groups, and
wherein the secondary building units are connected to the organic ligands through the oxygen atoms of the carboxylate groups in the organic ligands.
2. The MOF of claim 1, wherein the secondary building units form 1D rod-like chains or distinct multinuclear metal clusters.
3. The MOF of claim 1 or 2, wherein the organic ligands comprise one or more moieties of Formula (I)-(XII):



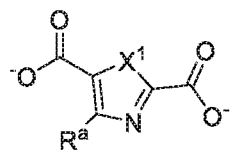
(I), wherein X¹ is NH, O or S, and each of R^a and R^b is independently H or alkyl;



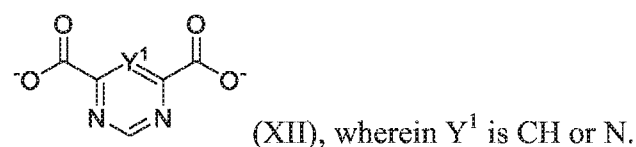
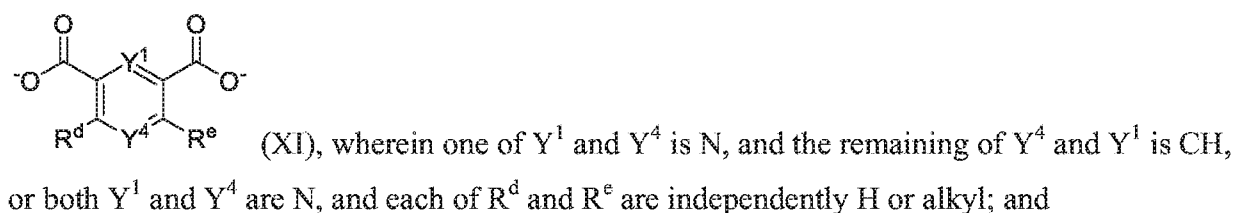
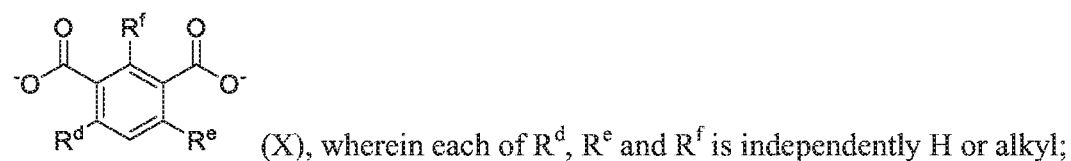
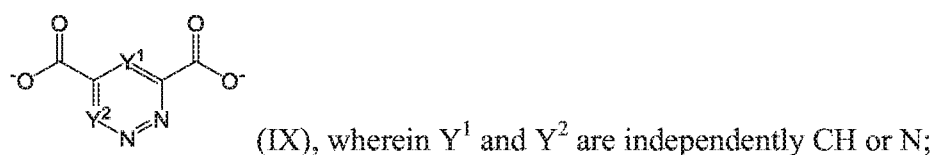
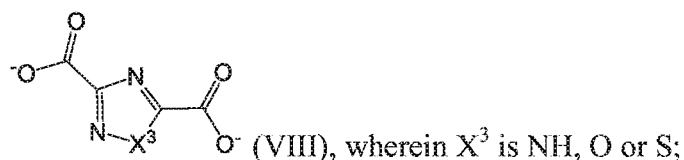
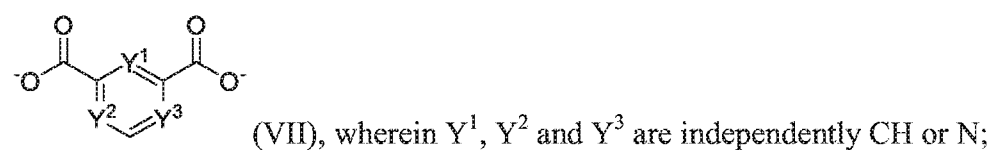
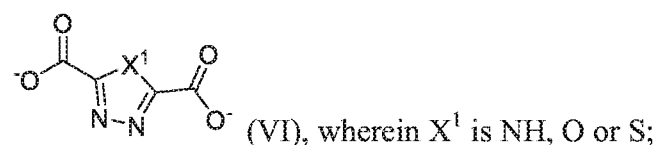
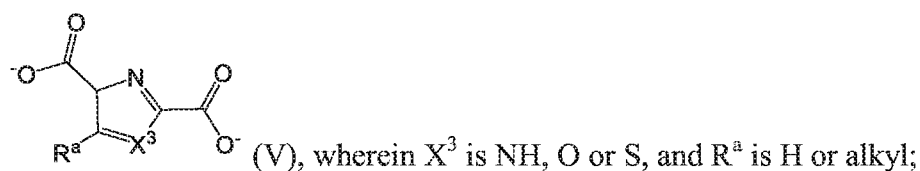
(II), wherein X³ is NH, O or S, and each of R^a and R^c is independently H or alkyl;



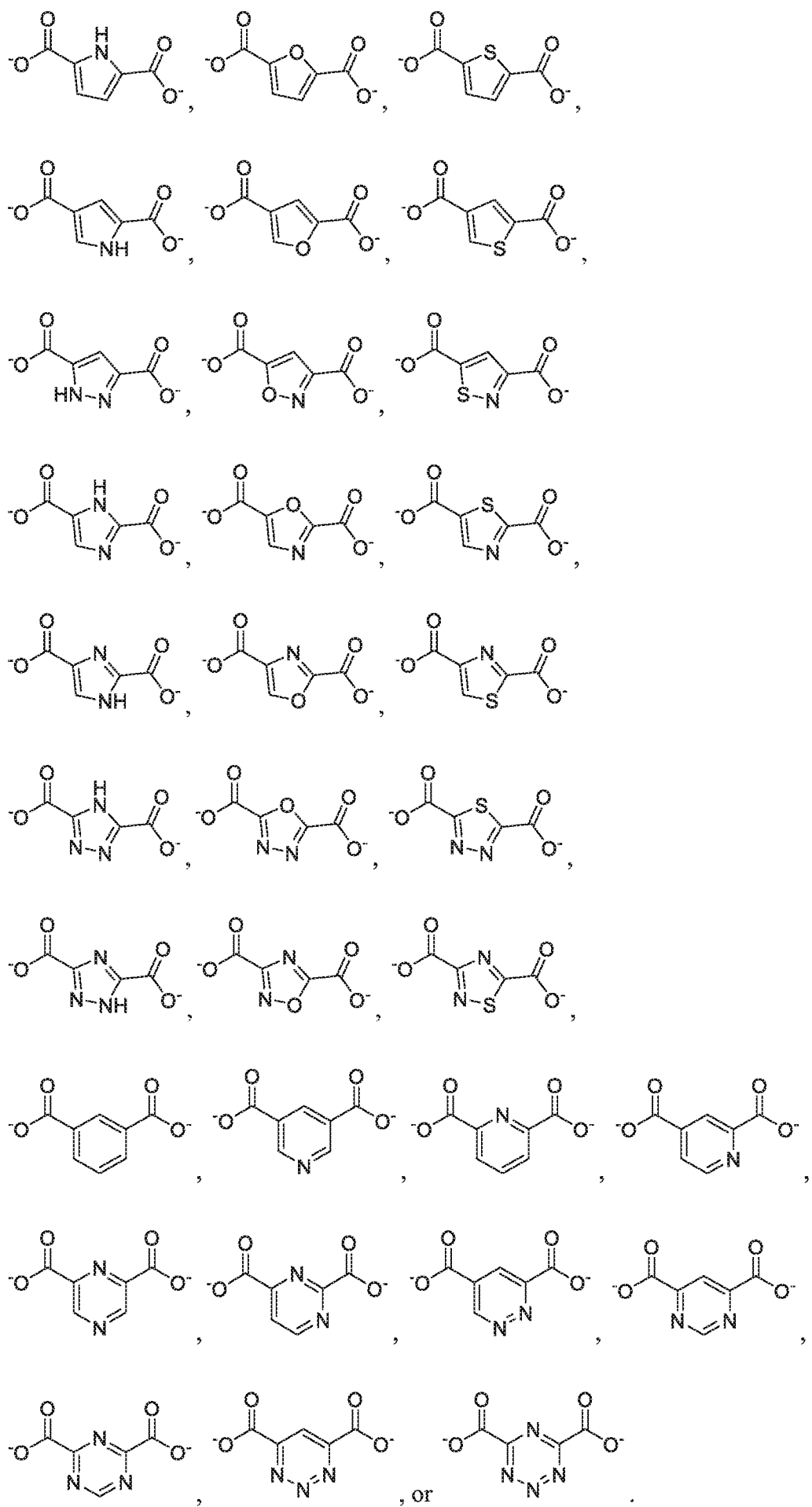
(III), wherein X² is NH, O or S, and R^c is H or alkyl;



(IV), wherein X¹ is NH, O or S, and R^a is H or alkyl;

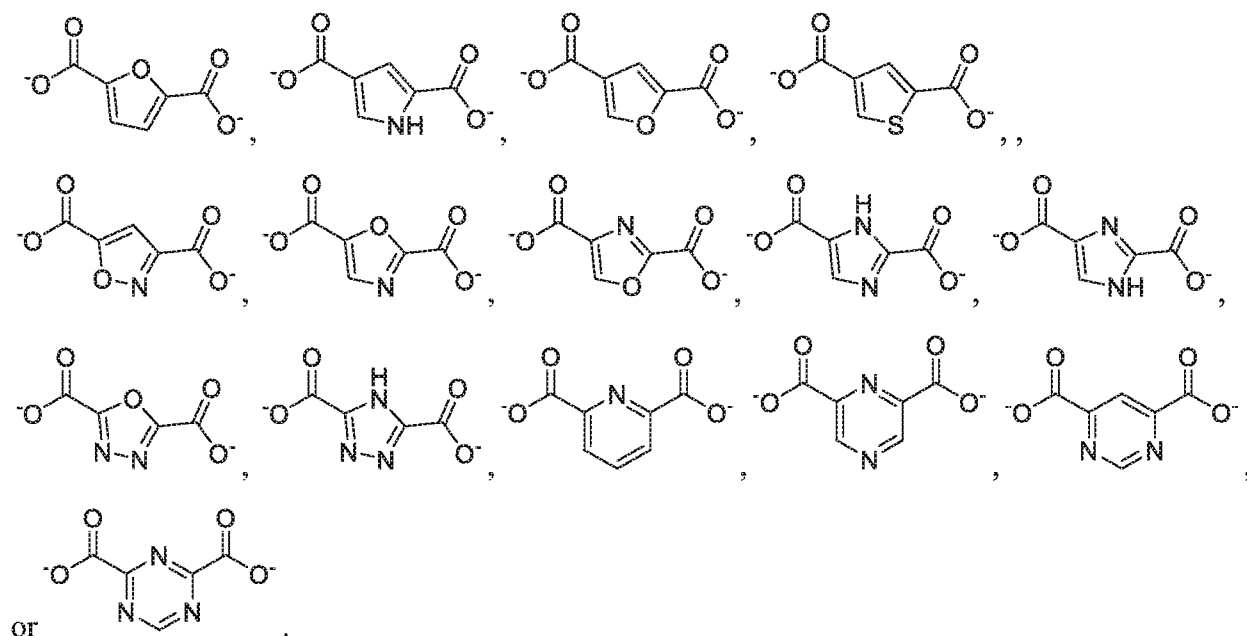


4. The MOF of claim 3, wherein R^a , R^b , R^c , R^d , R^e and R^f are each H.
5. The MOF of any one of claims 1 to 4, wherein the organic ligands comprise one or more moieties:



6. The MOF of any one of claims 1 to 5, wherein the MOF is a single-linker metal-organic framework.

7. The MOF of claim 6, wherein each organic ligand comprises:



8. The MOF of any one of claims 1 to 7, wherein each secondary building unit comprises aluminum hydroxide.

9. The MOF of any one of claims 1 to 7, wherein the MOF is MOF-313, MOF-314, or MOF-323.

10. The MOF of any one of claims 1 to 5, wherein the MOF is a multivariate metal-organic framework (MTV-MOF).

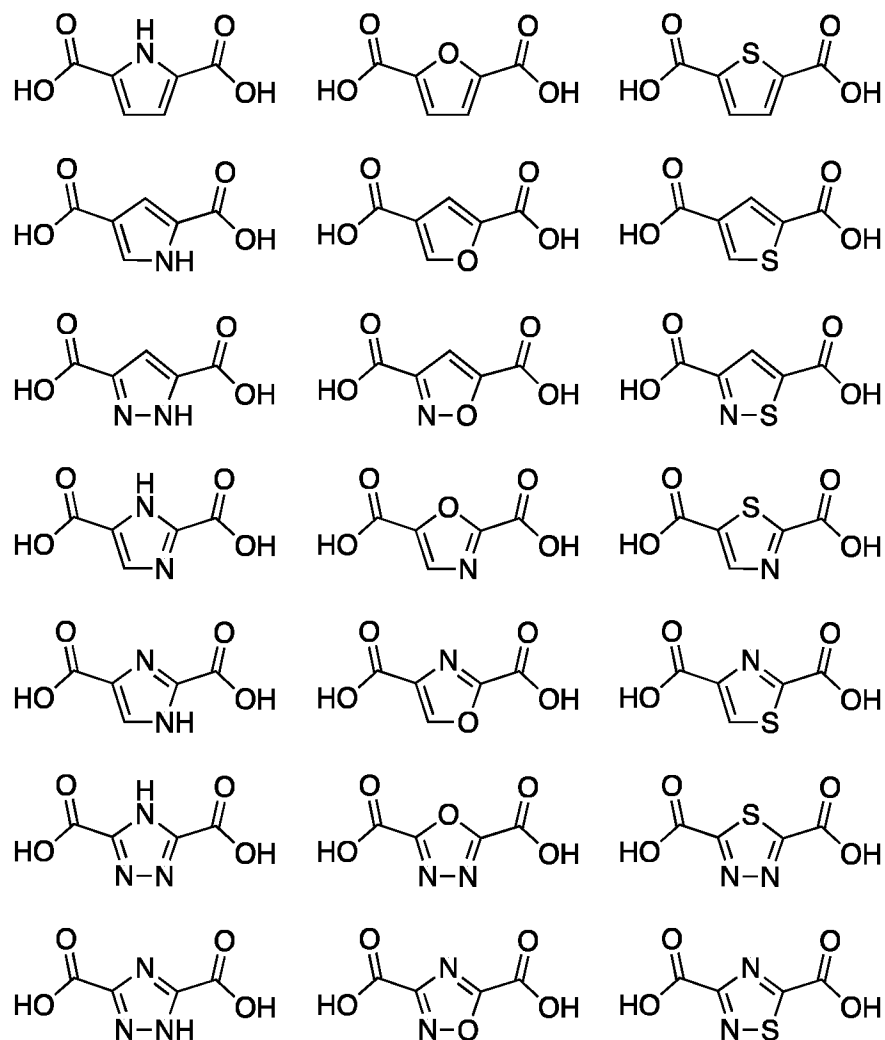
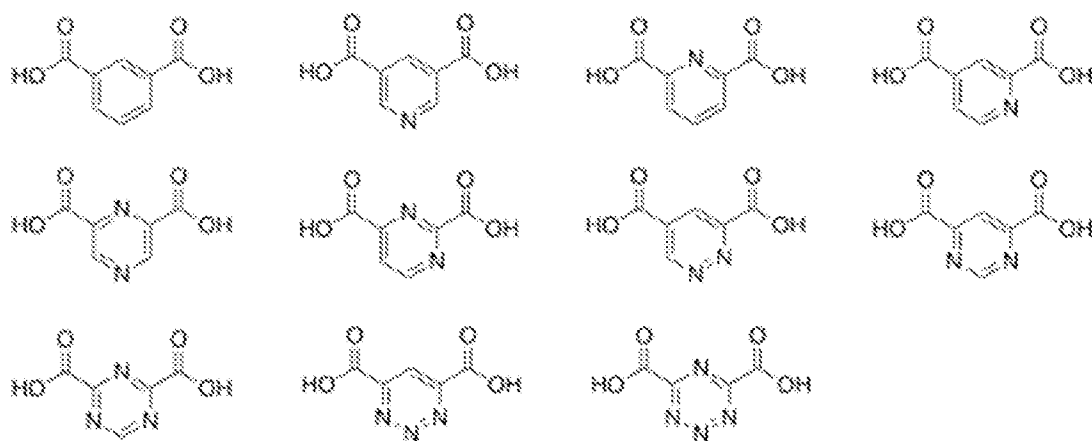
11. The MOF of claim 10, wherein the organic ligands comprises a plurality of moieties of Formula (I)-(XII).

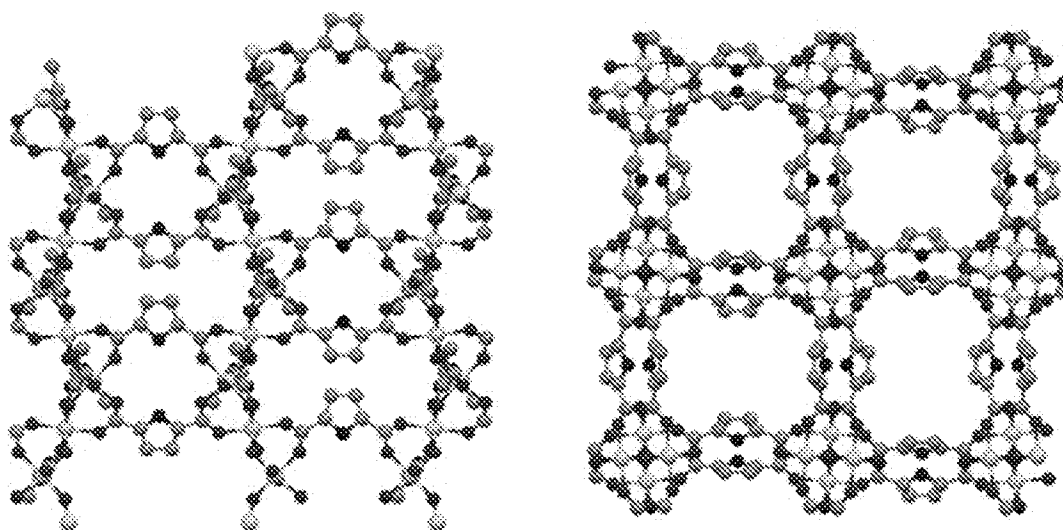
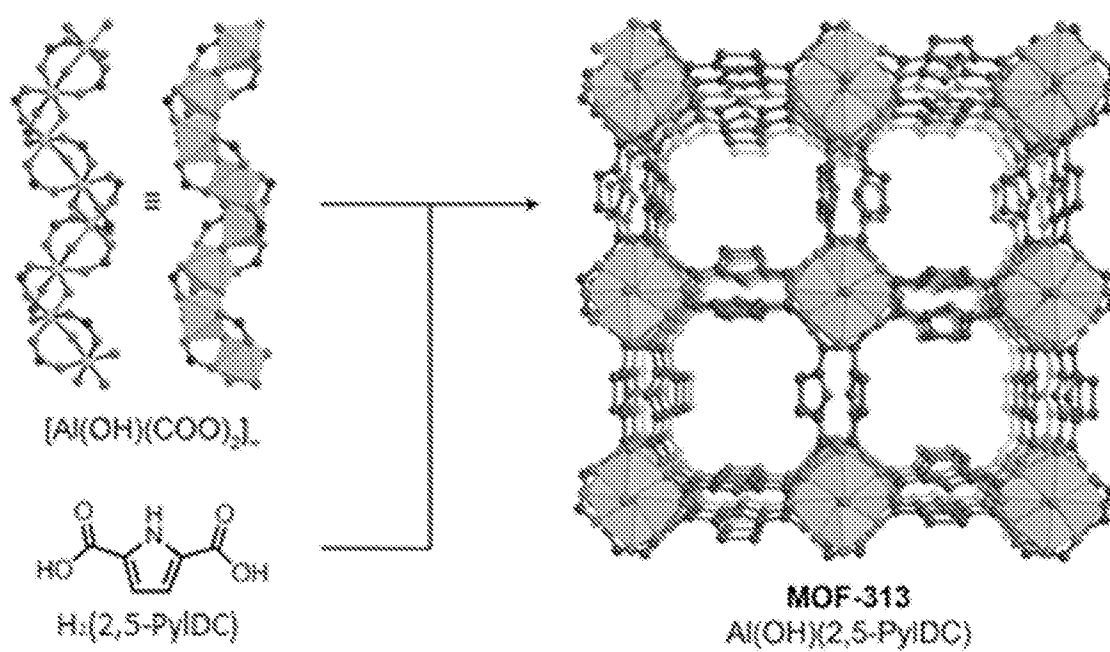
12. The MOF of claim 11, wherein the organic ligands comprises two moieties of Formula (I)-(XII).

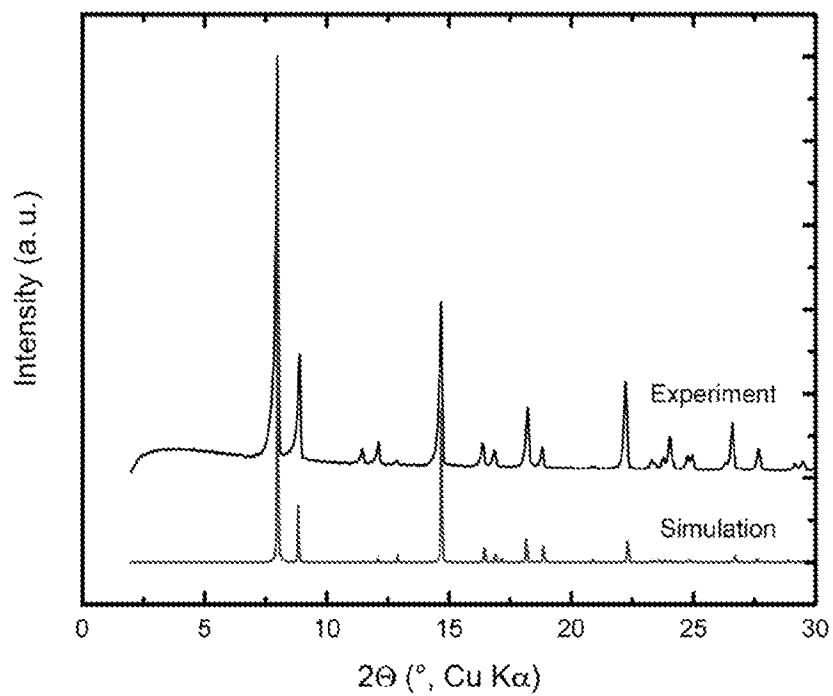
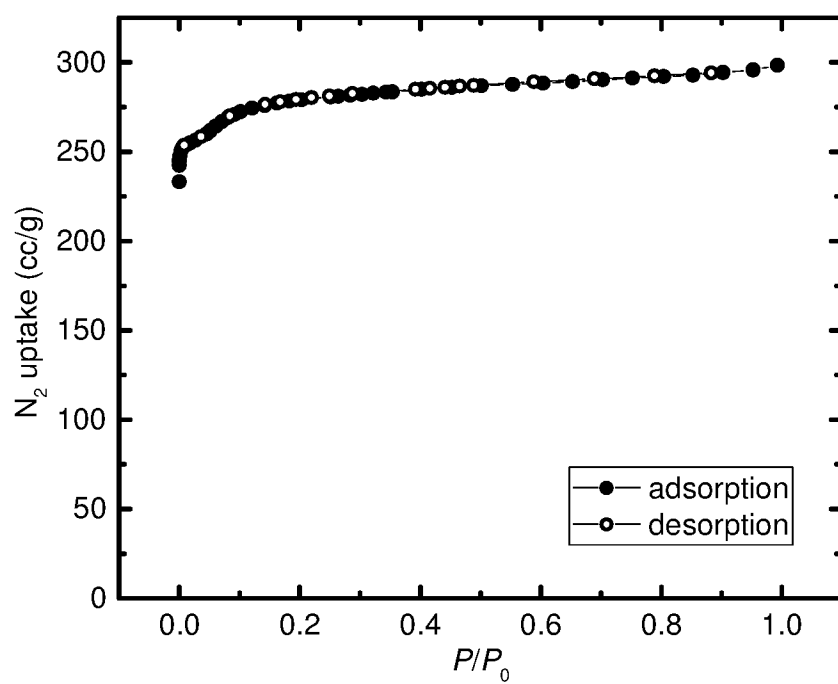
13. The MOF of any one of claims 1 to 5, wherein the MOF is $\text{Al}(\text{OH})(2,5\text{-FDC})_w(2,4\text{-TDC})_v$, wherein $w + v = 1$; or $\text{Al}(\text{OH})(3,5\text{-PynDC})_m(\text{IPA})_n$, wherein $m + n = 1$.

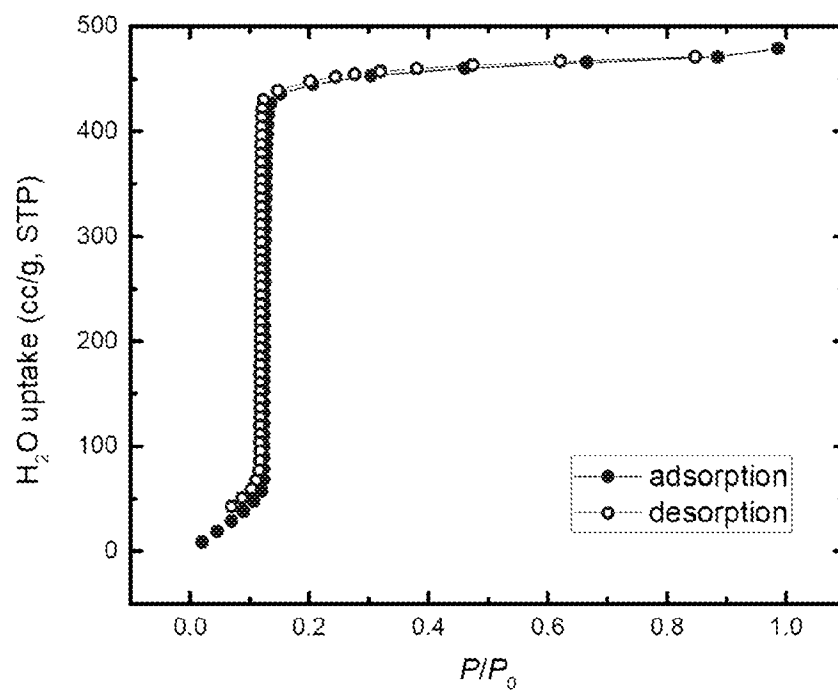
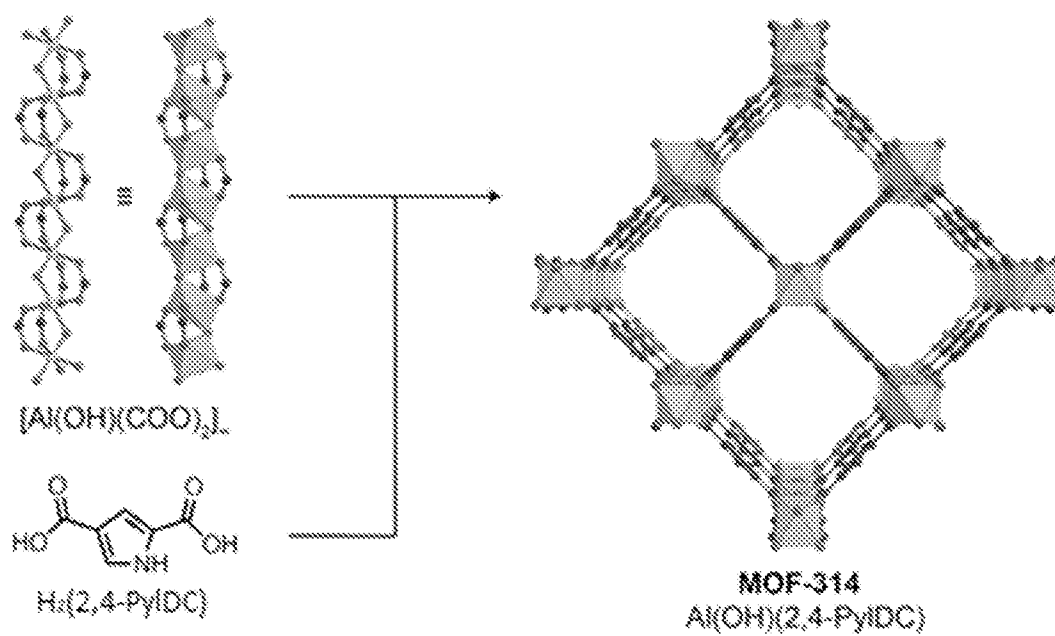
14. The MOF of any one of claims 1 to 13, wherein each secondary building unit comprises zirconium, nickel, iron, copper, manganese, aluminum, magnesium, calcium, strontium, barium, titanium, zinc, indium, cadmium, hafnium, lead, cobalt, or chromium, or a complex thereof.

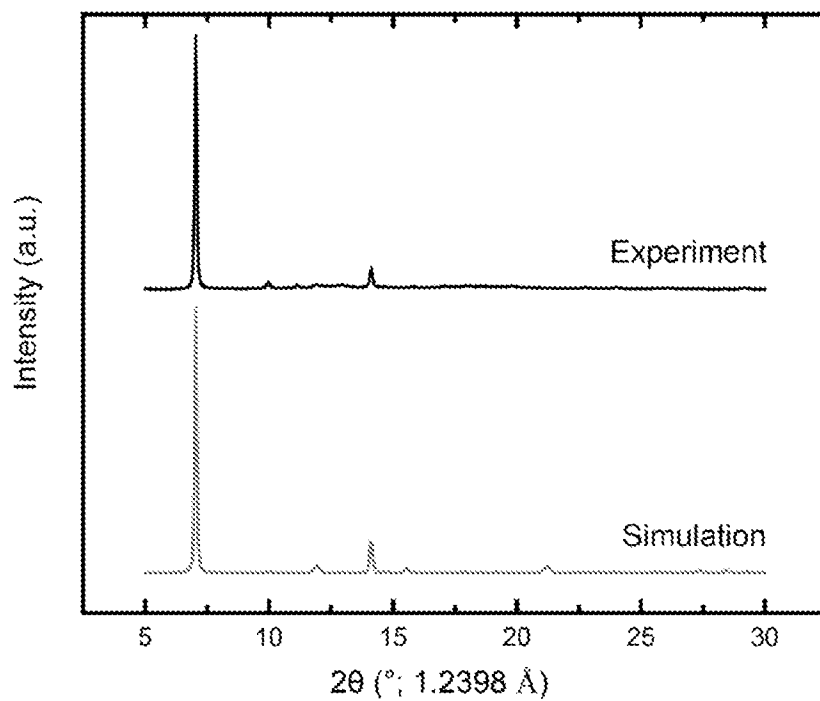
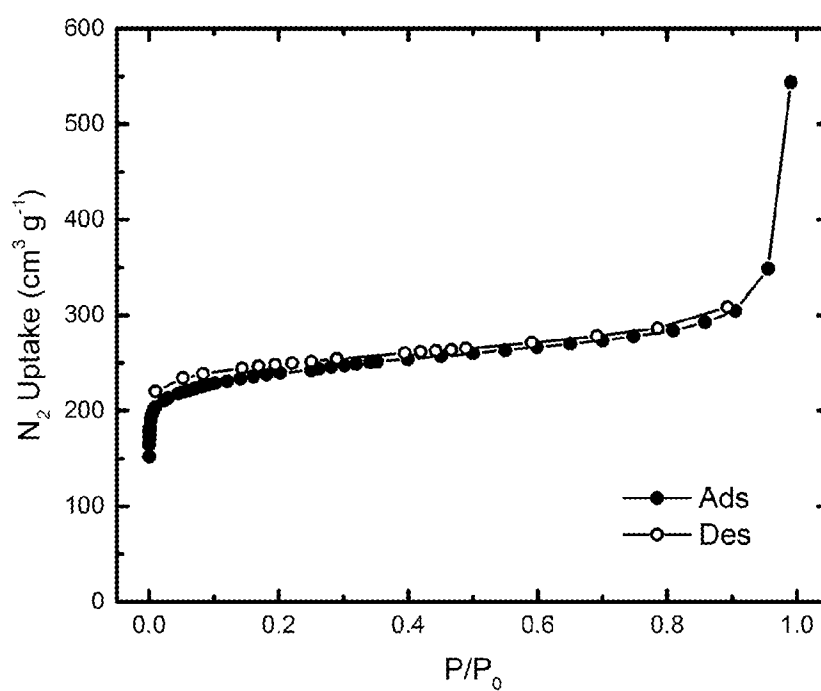
15. The MOF of any one of claims 1 to 13, wherein each secondary building unit comprises aluminum or an aluminum-containing complex.
16. The MOF of any one of claims 1 to 15, wherein the secondary building units form 1D rod-like chains, and the organic ligands comprise one or more v-shaped ditopic moieties.
17. The MOF of claim 16, wherein the secondary building units are connected to the organic ligands through the oxygen atoms of the carboxylate groups in the organic ligands via a cis-edge-shared octahedra geometry or a trans-edge-shared octahedra geometry.
18. A method of water harvesting, comprising:
 - adsorbing water from ambient atmosphere using a water-harvesting system, wherein the water-harvesting system comprises one or more MOFs of any one of claims 1 to 17;
 - desorbing vapor from the one or more MOFs; and
 - collecting water from the vapor.
19. The method of claim 18, wherein the water is collected from the vapor using a condenser.
20. A water harvesting system, comprising an adsorbent layer comprising one or more MOFs of any one of claims 1 to 17.
21. The system of claim 20, further comprising a condenser adjacent to the adsorbent layer.
22. The system of claim 20, wherein the system is a passive water harvesting system.
23. A MOF of any one of claims 1 to 17 for use in water harvesting or water purification.
24. A MOF of any one of claims 1 to 16 for use in heat pumps, dehumidifiers, adsorption refrigerators, solar cooling systems, dryers, organic light emitting devices, or secondary battery devices.
25. A method of using the MOF of any one of claims 1 to 16, comprising containing in, storing in and/or extracting from the MOF a predetermined gas or fluid.
26. The method of claim 24, wherein the gas or fluid is CO₂, H₂O, H₂, CH₄, C₂H₄, or C₂H₂, or any combinations thereof.

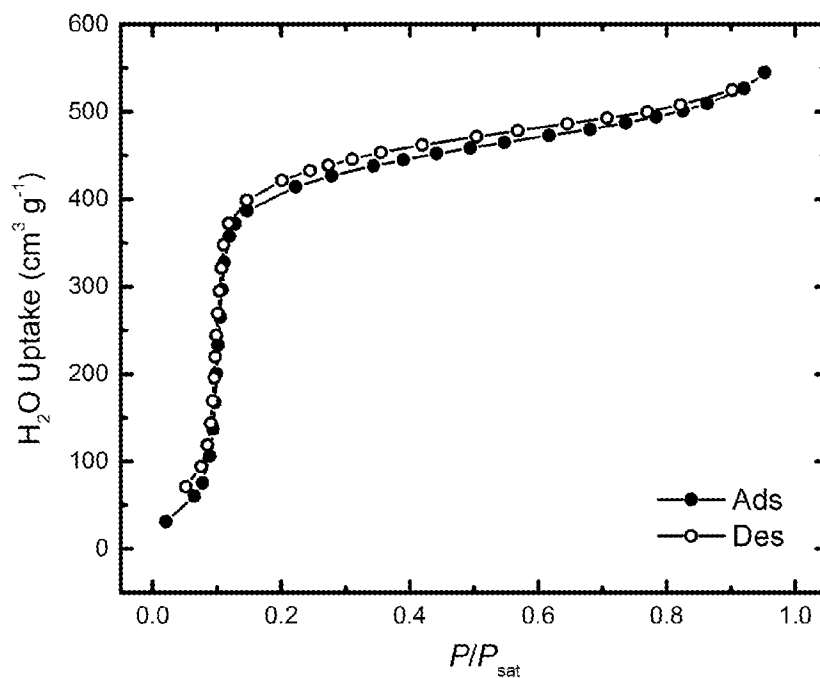
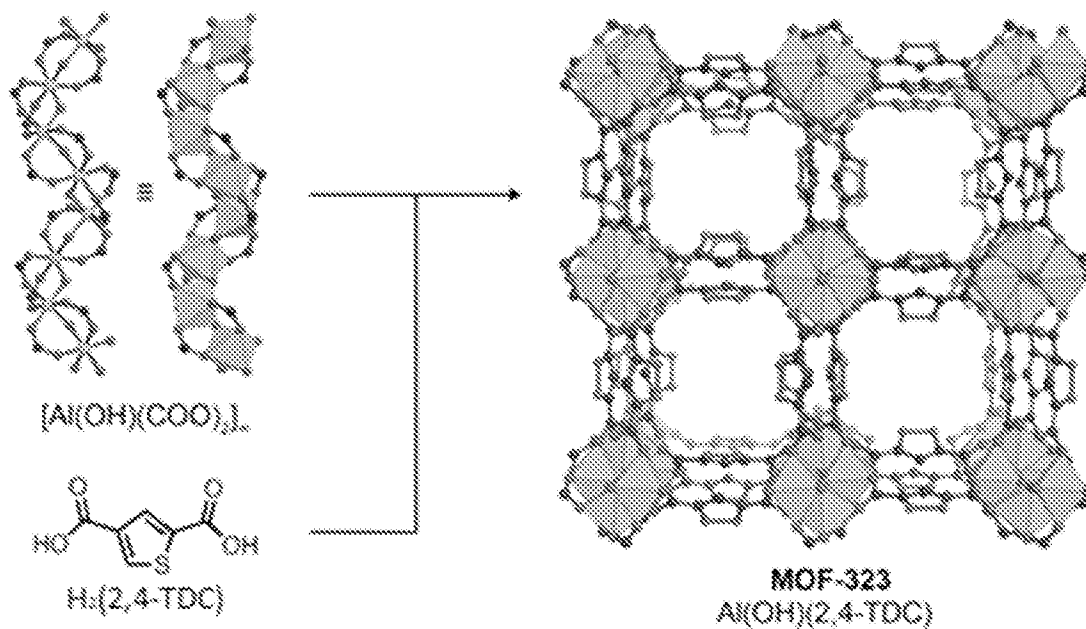
**FIG. 1A****FIG. 1B**

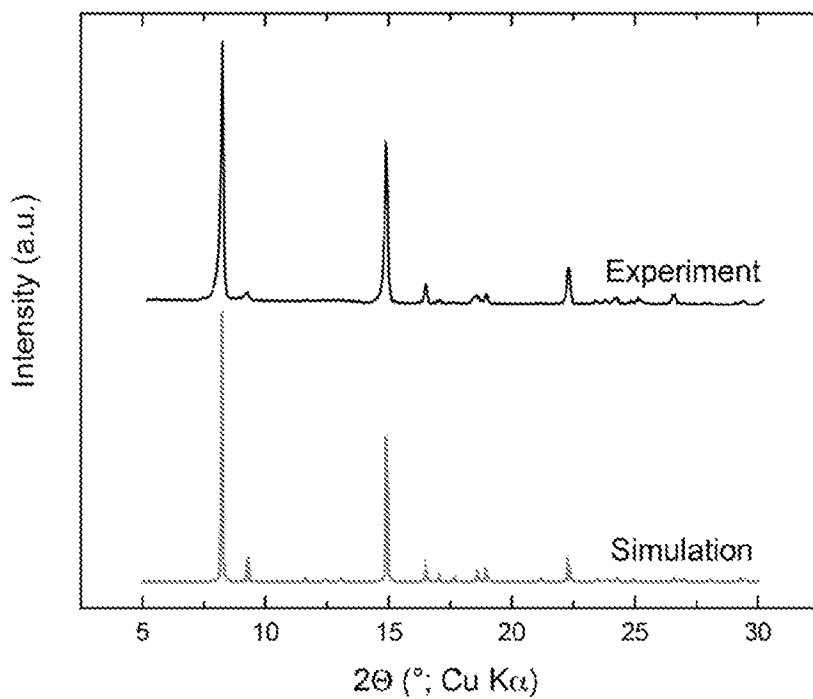
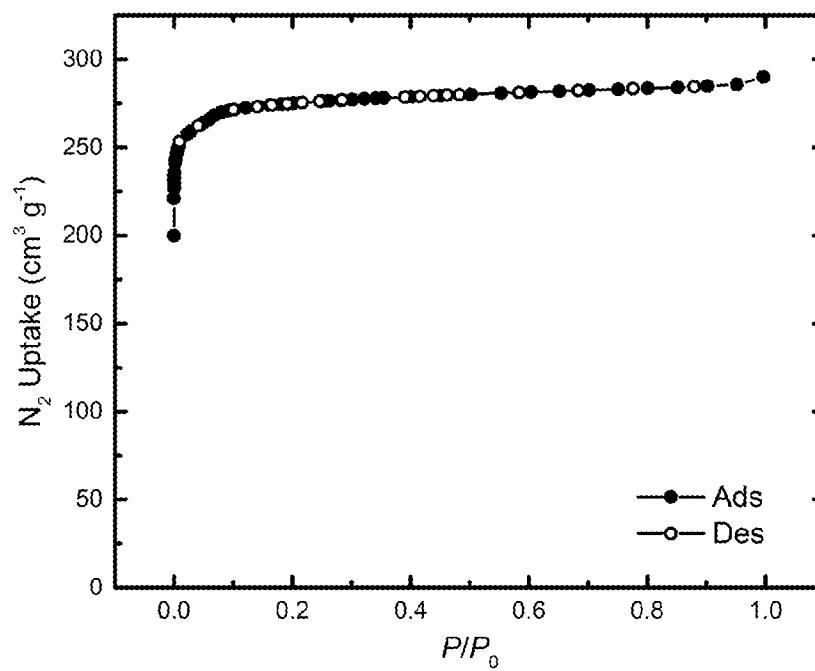
**FIG. 2****FIG. 3**

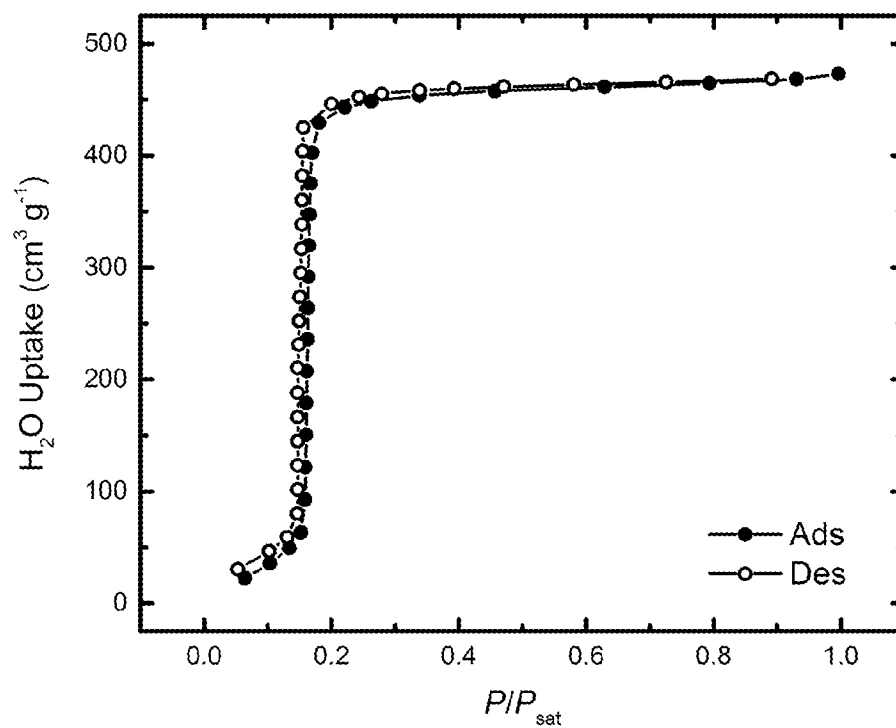
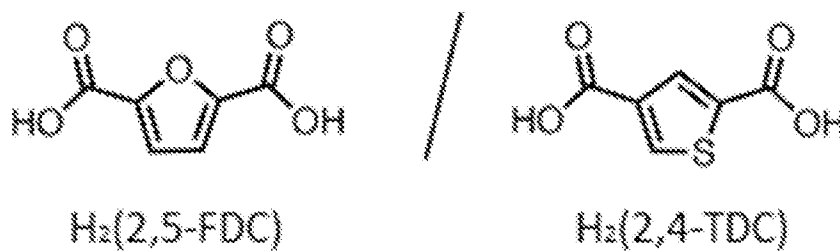
**FIG. 4****FIG. 5**

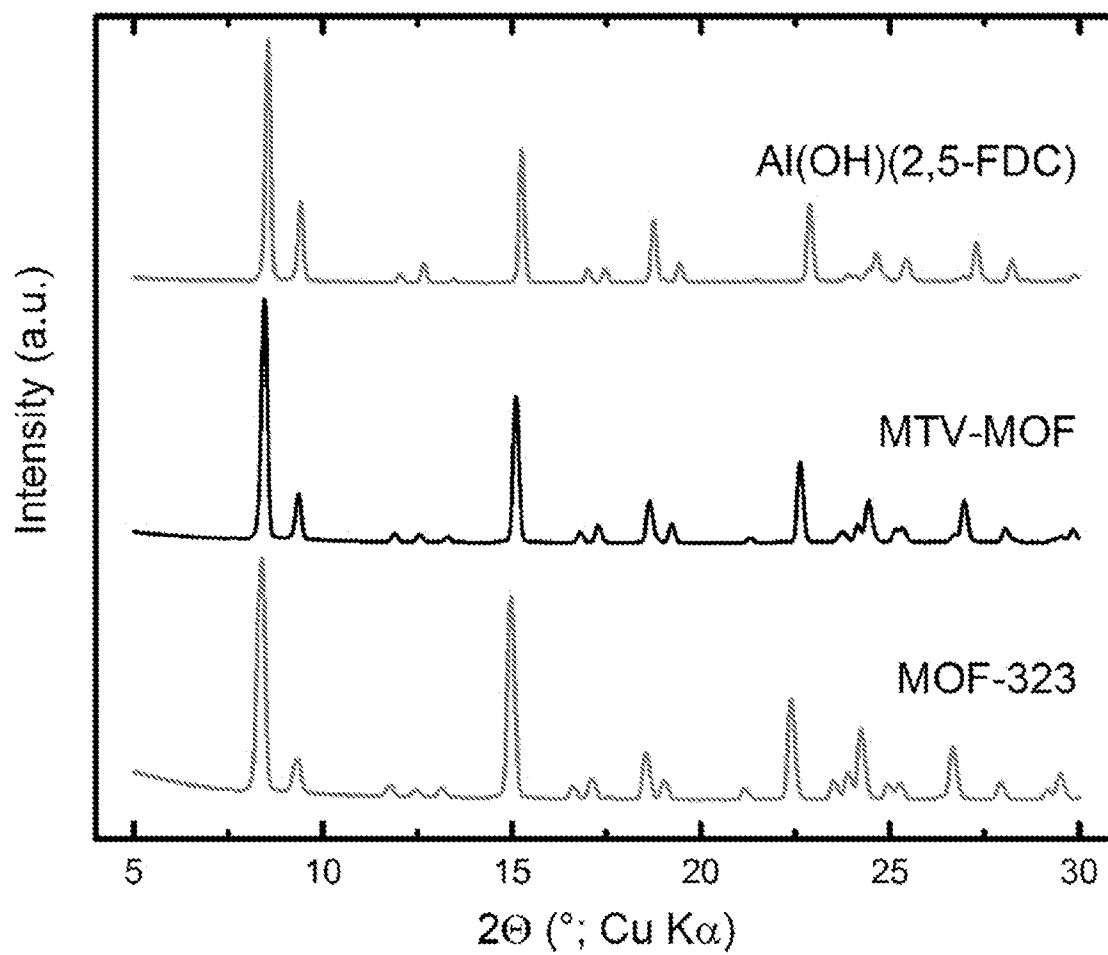
**FIG. 6****FIG. 7**

**FIG. 8****FIG. 9**

**FIG. 10****FIG. 11**

**FIG. 12****FIG. 13**

**FIG. 14****FIG. 15**

***FIG. 16***

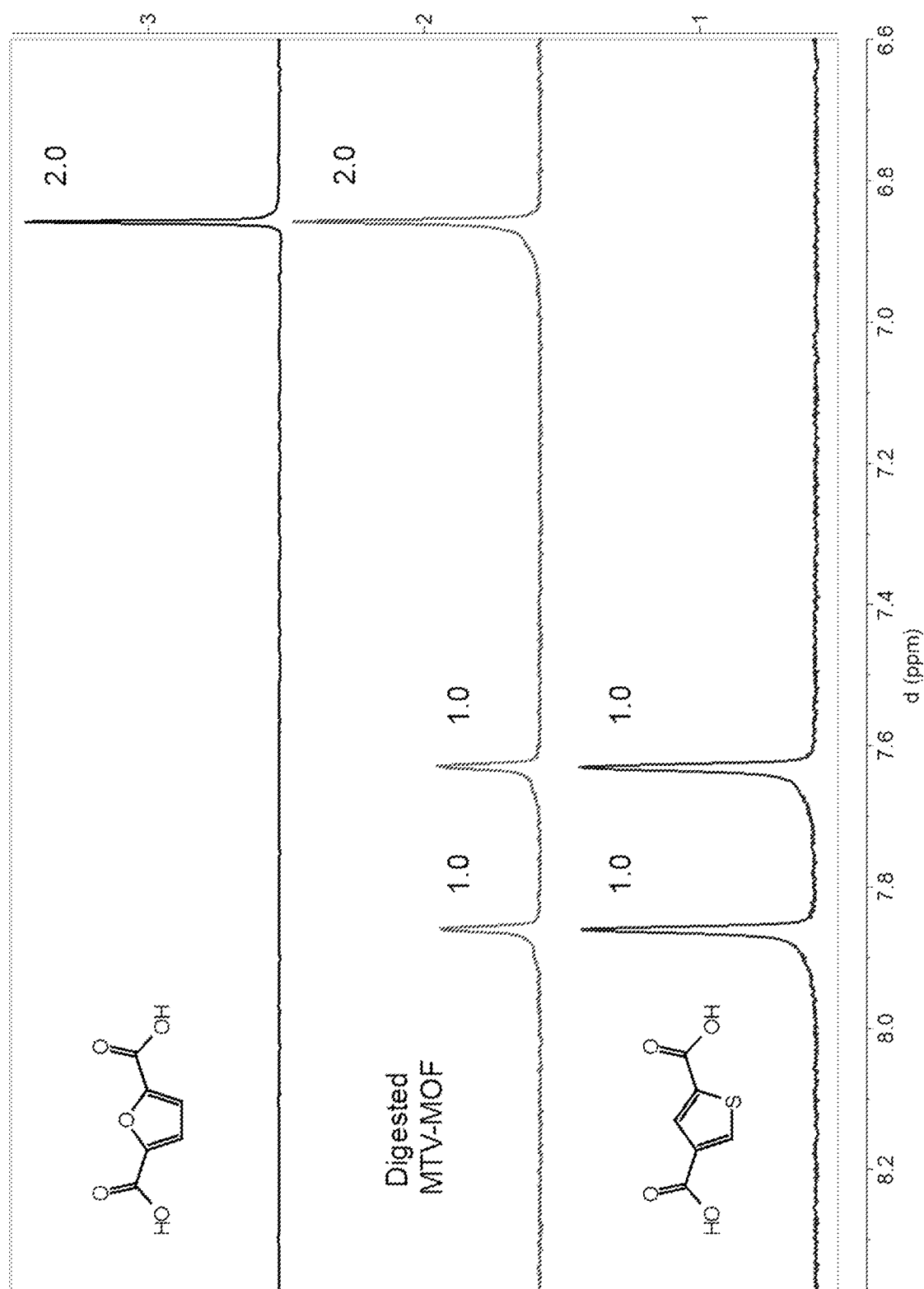


FIG. 17

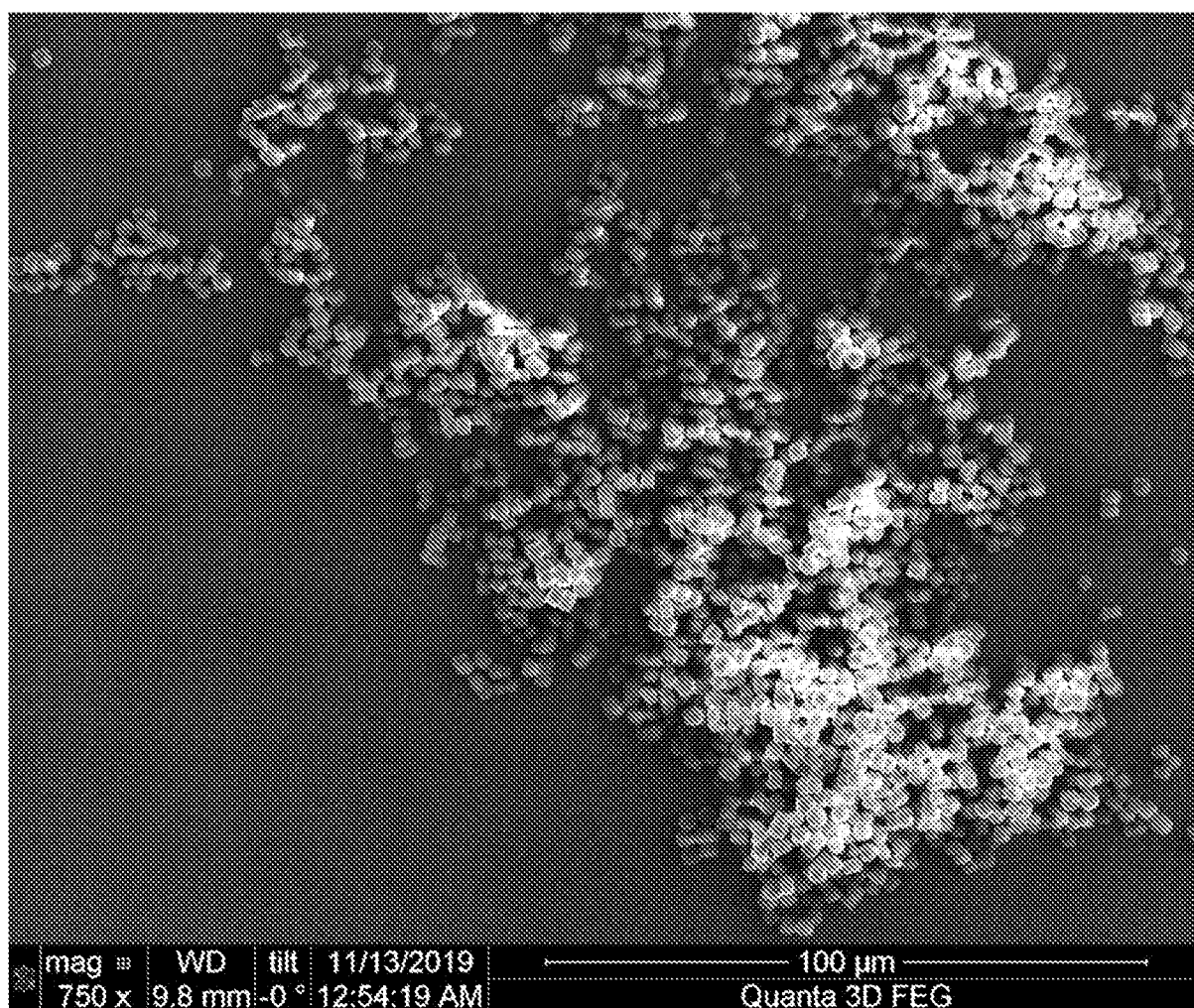


FIG. 18

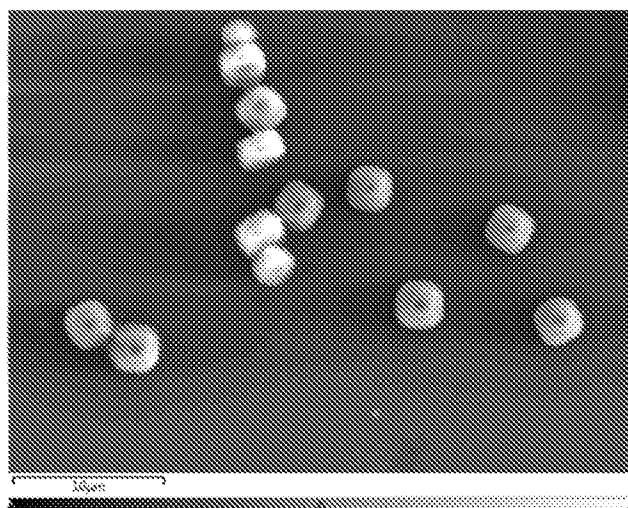


FIG. 19A

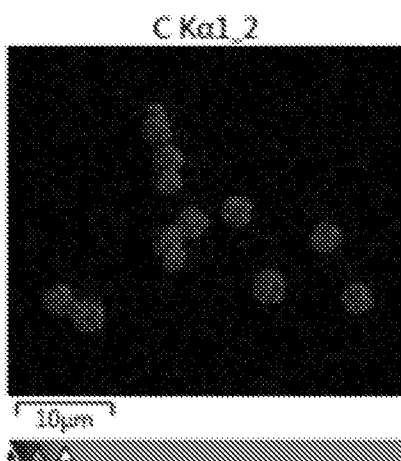


FIG. 19B

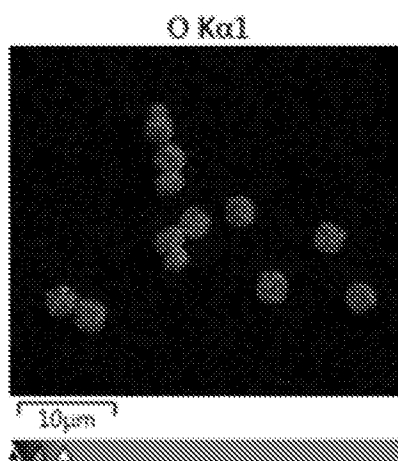


FIG. 19C

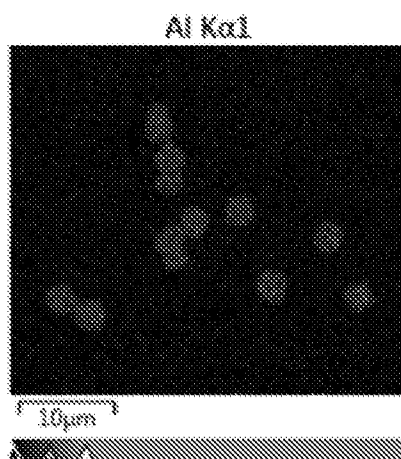


FIG. 19D

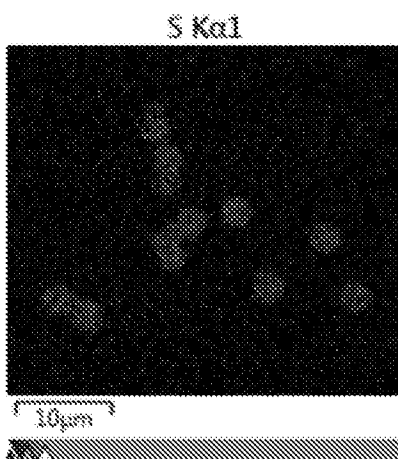
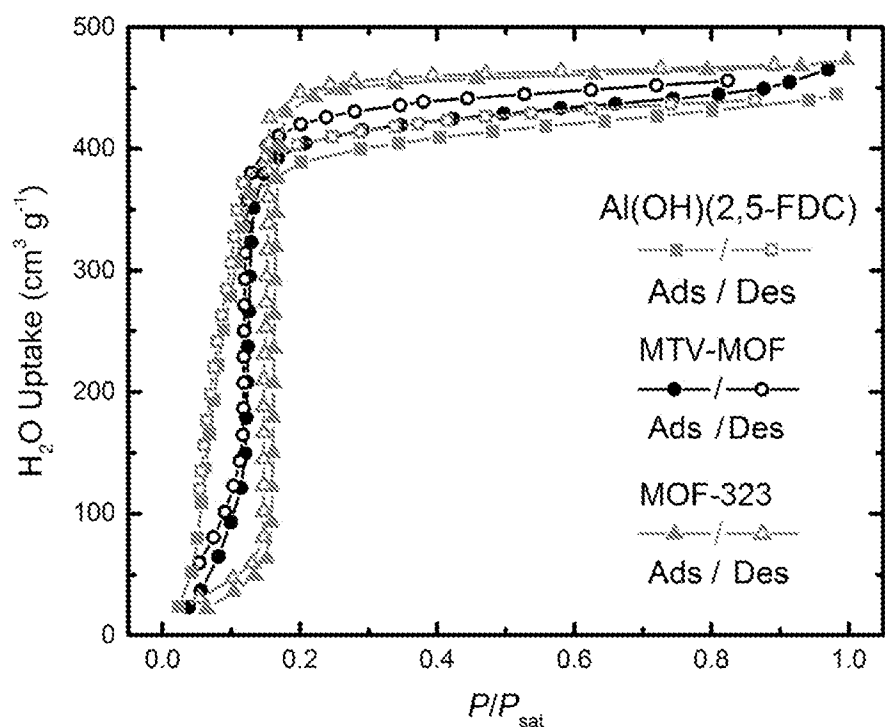
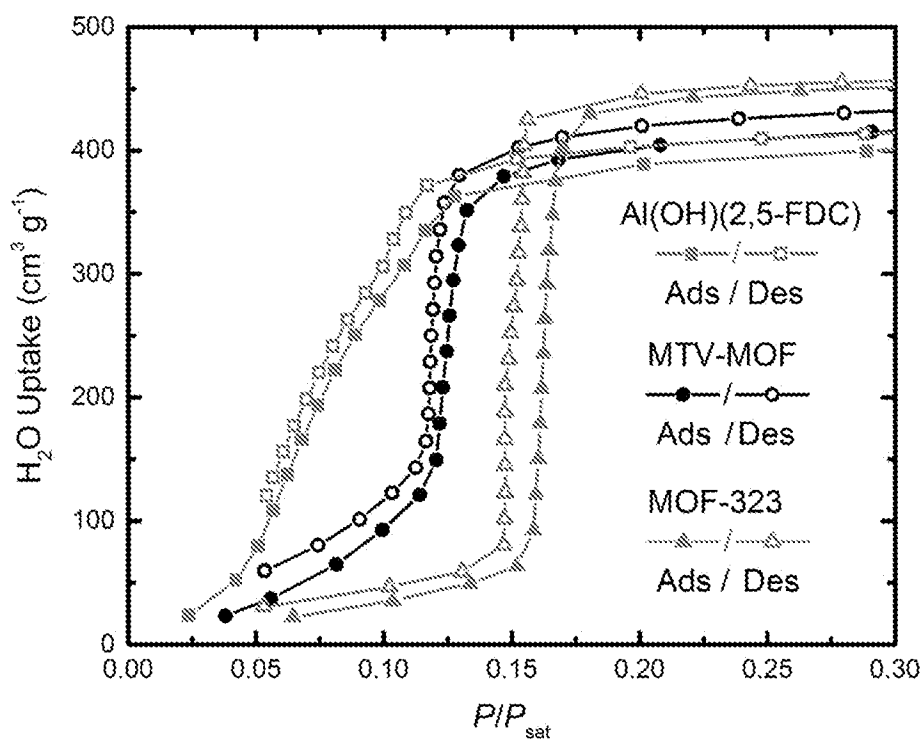
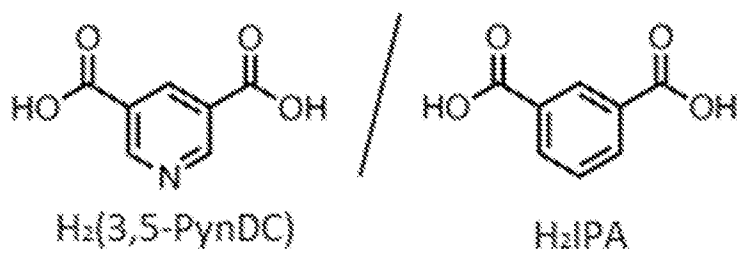
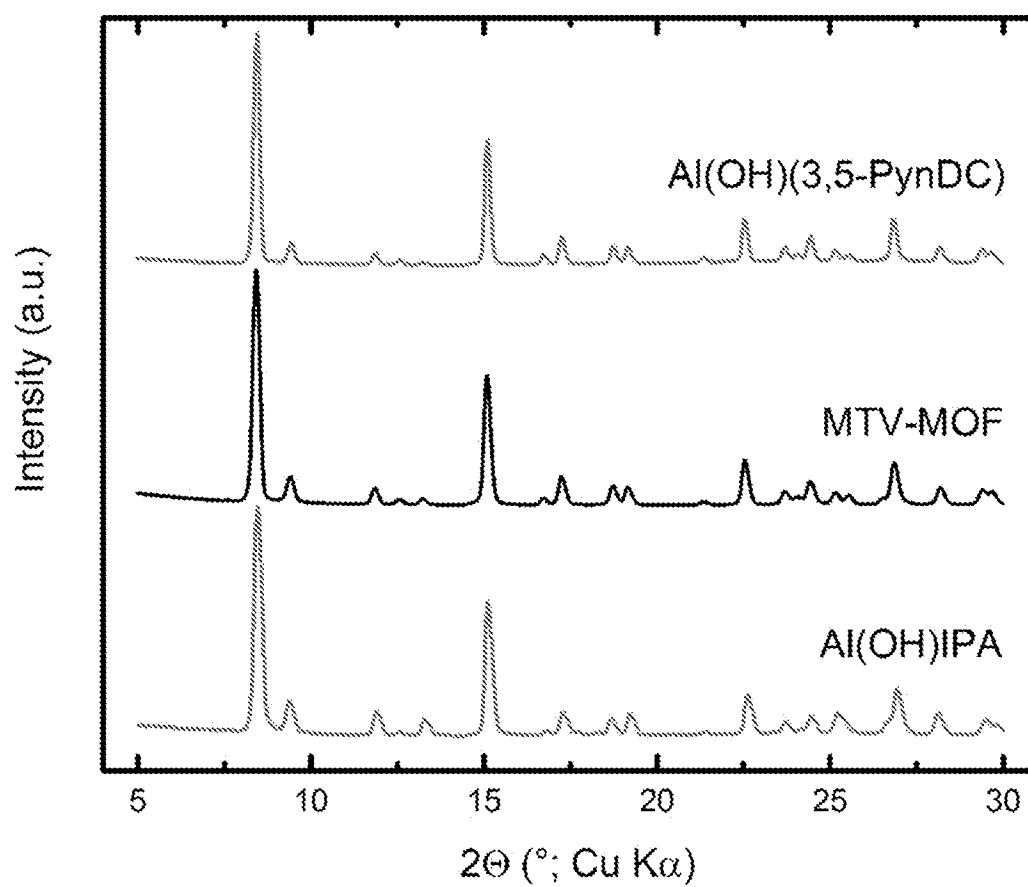


FIG. 19E

**FIG. 20****FIG. 21**

**FIG. 22****FIG. 23**

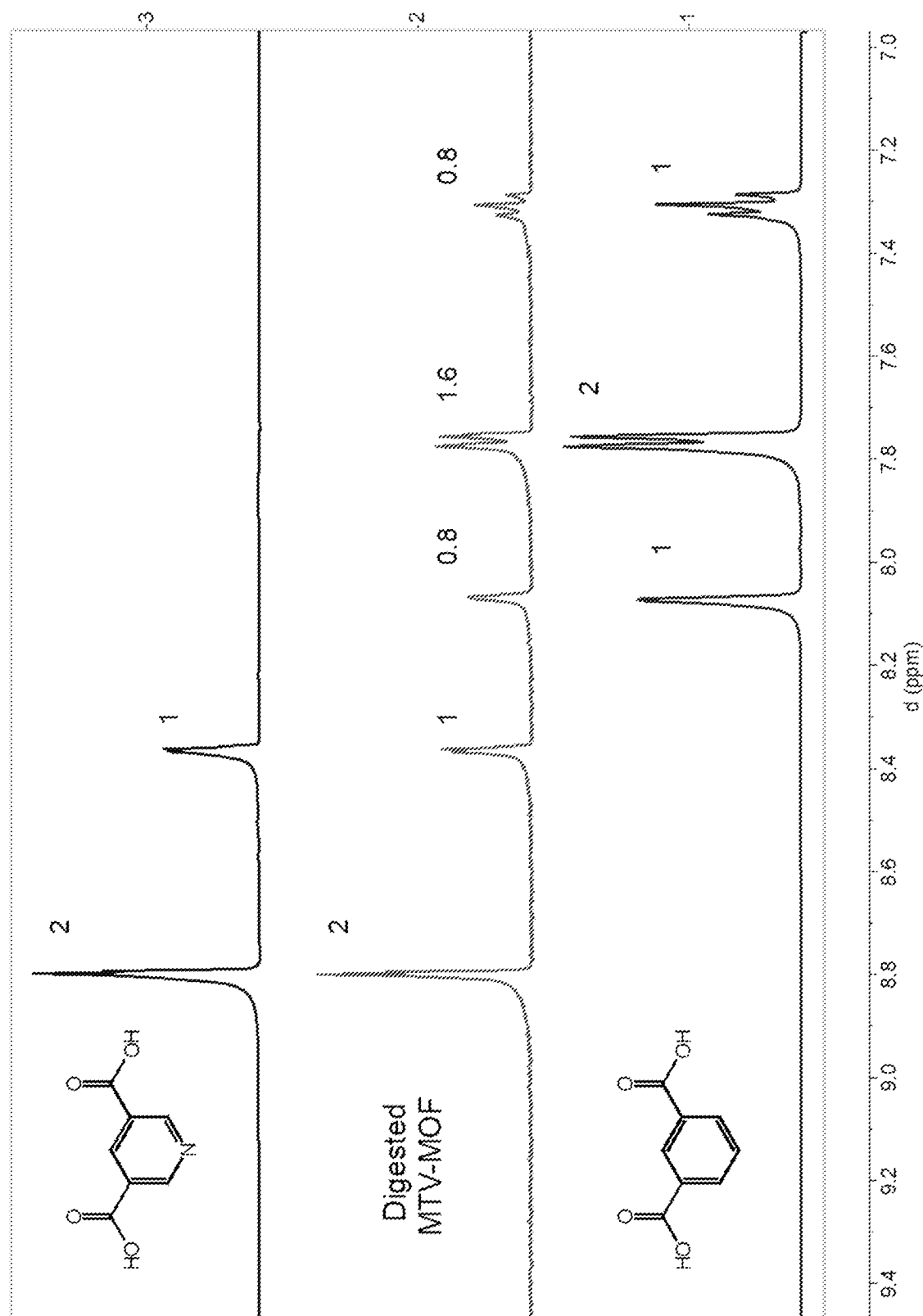


FIG. 24

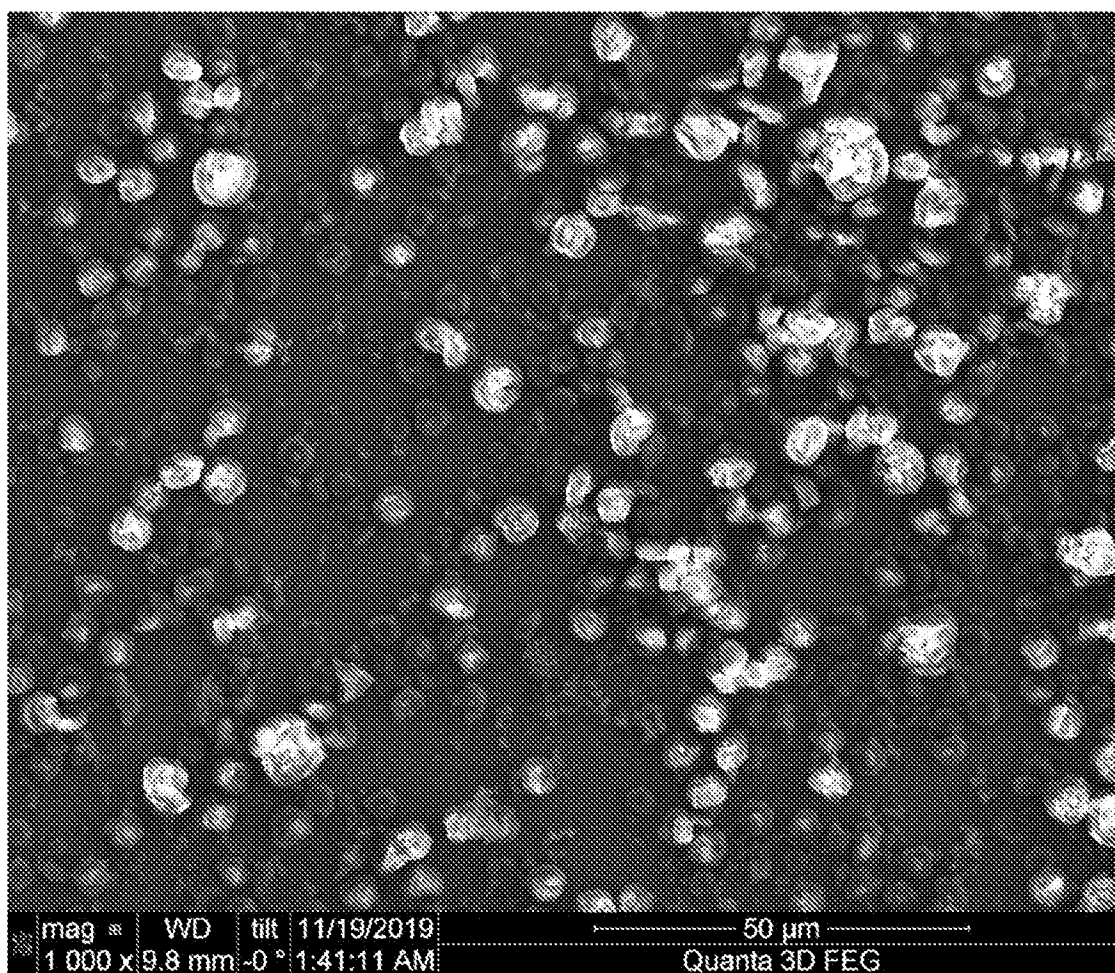


FIG. 25

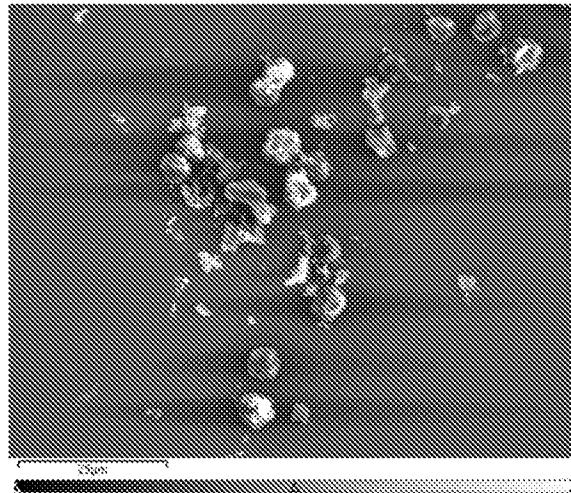


FIG. 26A

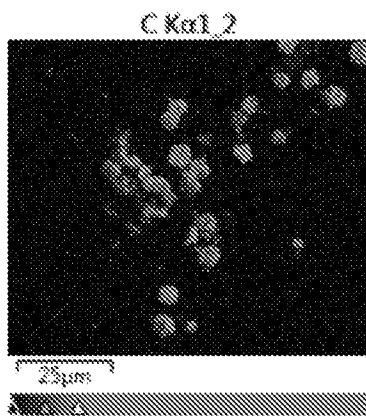


FIG. 26B

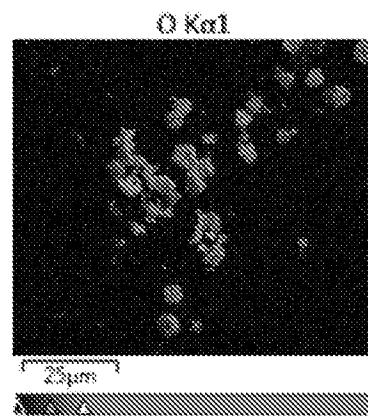


FIG. 26C

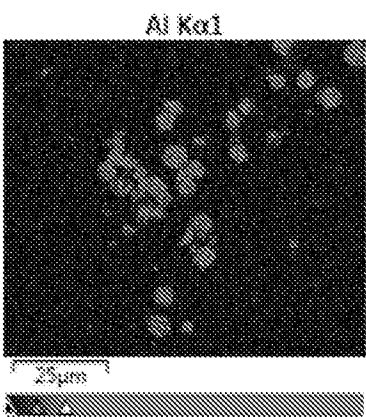


FIG. 26D

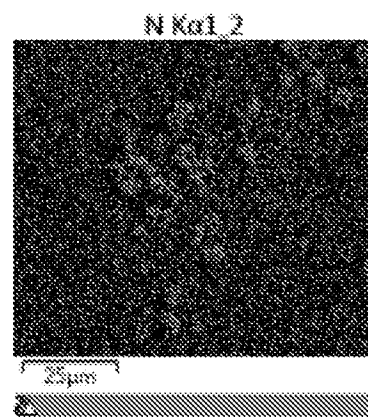


FIG. 26E

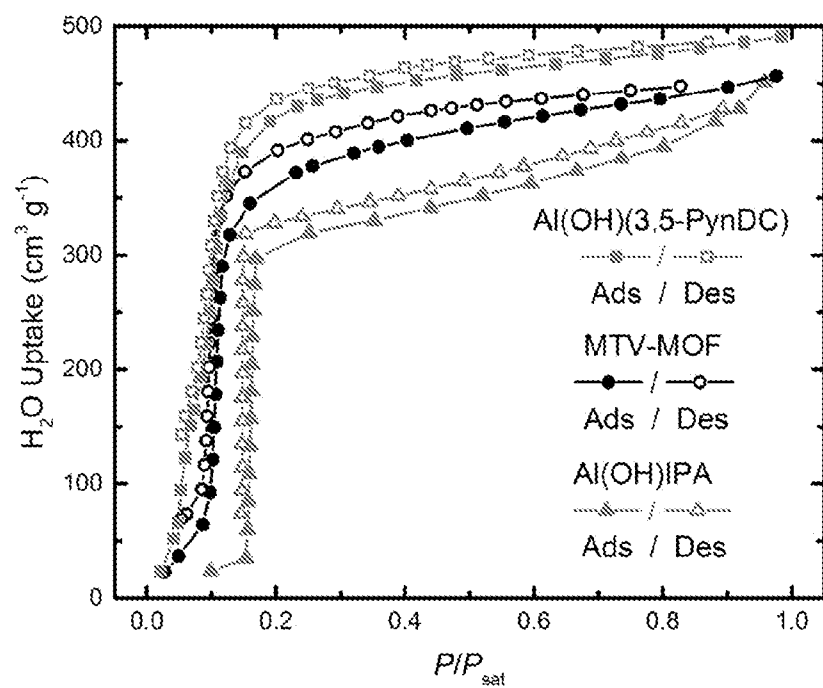


FIG. 27

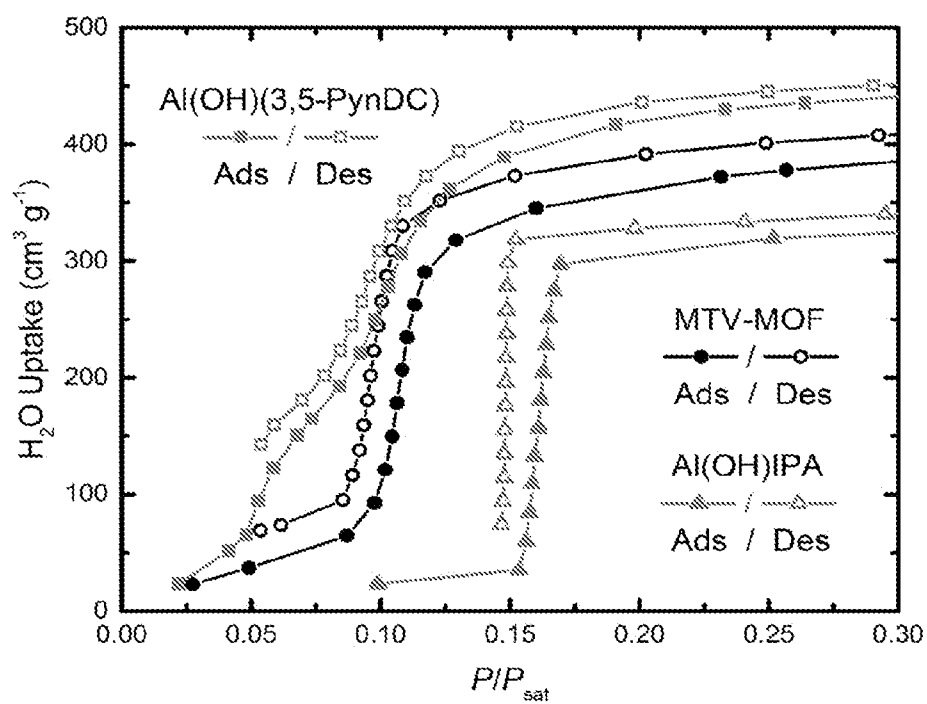


FIG. 28

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US19/63442

A. CLASSIFICATION OF SUBJECT MATTER

IPC - B01D 53/22; B01J 20/22, 31/16 (2020.01)

CPC - B01D 53/22; B01J 20/22, 31/16

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

See Search History document

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

See Search History document

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

See Search History document

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 2017/0008915 A1 (THE REGENTS OF THE UNIVERSITY OF CALIFORNIA) 12 January 2017; paragraphs [0005], [0009]	1-2, 3/1-2, 4/3/1-2
A	US 2012/0172612 A1 (YAGHI, OM et al.) 05 July 2012; entire document	1-2, 3/1-2, 4/3/1-2
A	US 2013/0061752 (FARHA, OK et al.) 14 March 2013; entire document	1-2, 3/1-2, 4/3/1-2



Further documents are listed in the continuation of Box C.



See patent family annex.

* Special categories of cited documents:

"A" document defining the general state of the art which is not considered to be of particular relevance

"D" document cited by the applicant in the international application

"E" earlier application or patent but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

"&" document member of the same patent family

Date of the actual completion of the international search

08 January 2020 (08.01.2020)

Date of mailing of the international search report

22 JAN 2020

Name and mailing address of the ISA/US

Mail Stop PCT, Attn: ISA/US, Commissioner for Patents
P.O. Box 1450, Alexandria, Virginia 22313-1450

Facsimile No. 571-273-8300

Authorized officer

Shane Thomas

Telephone No. PCT Helpdesk: 571-272-4300

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US19/63442

Box No. II Observations where certain claims were found unsearchable (Continuation of item 2 of first sheet)

This international search report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☐ Claims Nos.:
because they relate to subject matter not required to be searched by this Authority, namely:
2. ☐ Claims Nos.:
because they relate to parts of the international application that do not comply with the prescribed requirements to such an extent that no meaningful international search can be carried out, specifically:
3. ☒ Claims Nos.: 5-26
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box No. III Observations where unity of invention is lacking (Continuation of item 3 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

1. ☐ As all required additional search fees were timely paid by the applicant, this international search report covers all searchable claims.
2. ☐ As all searchable claims could be searched without effort justifying additional fees, this Authority did not invite payment of additional fees.
3. ☐ As only some of the required additional search fees were timely paid by the applicant, this international search report covers only those claims for which fees were paid, specifically claims Nos.:
4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this international search report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest and, where applicable, the payment of a protest fee.
- ☐ The additional search fees were accompanied by the applicant's protest but the applicable protest fee was not paid within the time limit specified in the invitation.
- ☐ No protest accompanied the payment of additional search fees.