



(86) **Date de dépôt PCT/PCT Filing Date:** 2017/02/03
(87) **Date publication PCT/PCT Publication Date:** 2017/08/31
(45) **Date de délivrance/Issue Date:** 2023/10/10
(85) **Entrée phase nationale/National Entry:** 2018/07/23
(86) **N° demande PCT/PCT Application No.:** JP 2017/003991
(87) **N° publication PCT/PCT Publication No.:** 2017/145708
(30) **Priorité/Priority:** 2016/02/26 (JP2016-036336)

(51) **Cl.Int./Int.Cl. H01M 8/18** (2006.01),
H01M 4/90 (2006.01)
(72) **Inventeurs/Inventors:**
KISHIMOTO, TAKEAKI, JP;
MIZUSHIRI, MAYUMI, JP
(73) **Propriétaire/Owner:**
NISSHINBO HOLDINGS INC., JP
(74) **Agent:** SMART & BIGGAR LP

(54) **Titre : CATALYSEUR CARBONE POUR ELECTRODES DE BATTERIE REDOX**
(54) **Title: CARBON CATALYST FOR REDOX FLOW BATTERY ELECTRODES**

(57) **Abrégé/Abstract:**

An electrode for redox flow batteries is produced using a carbon catalyst for redox flow battery electrodes, wherein a ratio of the number of oxygen atoms to the number of carbon atoms (O/C ratio) is 0.05 to 0.20 as measured by surface analysis using X-ray photoelectron spectroscopy



ABSTRACT

An electrode for redox flow batteries is produced using a carbon catalyst for redox flow battery electrodes, wherein a ratio of the number of oxygen atoms to the number of carbon atoms (O/C ratio) is 0.05 to 0.20 as measured by surface analysis using X-ray photoelectron spectroscopy.

DESCRIPTION

Title of Invention

CARBON CATALYST FOR REDOX FLOW BATTERY ELECTRODES

Technical Field

[0001]

The present invention relates to a catalyst for use in electrodes of redox flow batteries and particularly relates to a carbon catalyst for redox flow battery electrodes that consists of a carbonaceous material and that is suitable for use in vanadium redox flow batteries.

Background Art

[0002]

In recent years, a shift toward renewable energy generation such as wind power generation and photovoltaic power generation has been promoted to establish a low carbon society. However, wind power generation and photovoltaic power generation have a problem in that the power output varies depending on the time of day, the season, and the weather. To meet a large-scale electricity demand, it is necessary to take a measure for leveling load of electricity demand against the power output variation.

Power storage using storage batteries have been attracting interest as one of the measures for electric load leveling. Among storage batteries, redox flow batteries are particularly promising as batteries for large-scale power storage due to their features such as being easy to increase in size, being highly safe and stable, and having a long charge-discharge cycle life.

[0003]

Redox flow batteries are flow batteries that are charged and discharged through ion oxidation-reduction reactions induced by pump circulation of electrolyte solutions. Fig. 1 shows the outline of the configuration of a typical redox flow battery. The redox flow battery shown in Fig. 1 includes a cell 1, a catholyte tank 2, and an anolyte tank 3. The cell 1 is divided into a cathode cell 1a and an anode cell 1b by an ion-permeable separation membrane 4. A cathode 5 is provided in the cathode cell 1a, while an anode 6 is provided in the anode cell 1b. The cathode cell 1a is connected to the catholyte tank 2 that supplies and

receives a catholyte, and the catholyte is circulated by a catholyte pump 7. The anode cell 1b is connected to the anolyte tank 3 that supplies and receives an anolyte, and the anolyte is circulated by an anolyte pump 8. In the redox flow battery, cations contained in the electrolyte solutions (the catholyte and anolyte) move between the cathode cell 1a and anode cell 1b through the separation membrane 4 and undergo oxidation-reduction reaction both at the cathode 5 and at the anode 6, which leads to a change in valence of the ions contained in the electrolyte solutions. In this way, charge and discharge are repeated.

[0004]

Redox flow batteries as described above have the disadvantage of having a lower energy density than other storage batteries. Thus, research and development aimed at improvement in energy density have been conducted, and various modifications or proposals related to electrodes have been made in terms of facilitating electrode reactions in redox flow batteries.

[0005]

For example, Patent Literature 1 proposes a carbonaceous electrode able to reduce the change over time in charge-discharge cycle, the electrode consisting of carbonaceous fibers in which the number of oxygen atoms bound on the carbon surface is 10 to 25% of the number of carbon atoms.

Additionally, Patent Literature 2 describes a carbon electrode material for vanadium redox flow batteries, the electrode material comprising vapor-grown carbon fibers having an average fiber diameter of 0.05 to 0.3 μm and an average aspect ratio of 10 to 500.

Further, electrode production using oxidized graphene, partially-reduced graphite oxide, nitrogen-doped carbon, a metal-supported carbon material or the like has also been reported.

[0006]

Other than improvements to electrodes themselves, the use of a phthalocyanine, platinum catalyst, or the like as an electrode catalyst has also been reported. Patent Literature 3 states that an electrode catalyst obtained by modifying the surface of a carbon electrode with a metal-containing carbon thin film through thermal decomposition of metal phthalocyanine has high catalytic activity for an electrode reaction of a redox flow battery.

Citation List

Patent Literature

[0007]

PTL 1: JP 05-234612 A

PTL 2: JP 2006-156029 A

PTL 3: JP 2015-115158 A

Summary of Invention

Technical Problem

[0008]

However, the electrode described in Patent Literature 1 has a problem in that the electrode will have a high specific resistance since this electrode is treated by air oxidation so that the number of bound oxygen atoms in the surface of the electrode would be 10 to 25% of the number of carbon atoms. Additionally, it is difficult to increase the carbon crystallinity; thus, the electrode cannot be said to have sufficient resistance to oxidation, particularly when used in a redox flow battery employing an electrolyte solution containing 1.5 mol/l or more of vanadium ions. With this electrode, the cell resistance increases and the energy efficiency significantly decreases as charge-discharge cycles are repeated.

Additionally, the electrode described in Patent Literature 2 is unsatisfactory in terms of the reversibility of oxidation-reduction reaction.

Other conventional electrodes are difficult to produce for practical use. As for metal-supported carbon materials etc., it cannot be said that the specific metal species and the specific method for supporting have been thoroughly investigated.

[0009]

Additionally, conventional electrode catalysts have insufficient catalytic activity and durability or are expensive, and are therefore not practically useful. Even the electrode catalyst described in Patent Literature 3 cannot be said to have satisfactory catalytic activity. In an example of Patent Literature 3, a carbon thin film-coated glassy carbon (GC) plate electrode consisting of a GC plate coated with a metal-containing carbon thin film is produced, and this catalyst is thought to be not practically useful on carbon felt or carbon fiber non-woven fabric which is suitable for use as an electrode for redox flow batteries.

[0010]

Therefore, an electrode that provides improvement in energy density, in

particular, an electrode that facilitates electrode reactions and that has excellent durability, is demanded from the viewpoint of full-fledged practical use and widespread utilization of redox flow batteries.

[0011]

The present invention has been made to solve the above problems, and an object of the present invention is to provide a carbon catalyst for redox flow battery electrodes that is able to facilitate electrode reactions in redox flow batteries and that has excellent durability.

Solution to Problem

[0012]

The present invention has been made with a focus on electrode catalysts to improve the energy density of redox flow batteries and is based on the finding that a carbon catalyst having specific surface characteristics is able to facilitate electrode reactions and provide high durability.

[0013]

That is, the present invention provides the following [1] to [5].

[1] A carbon catalyst for redox flow battery electrodes, wherein a ratio of the number of oxygen atoms to the number of carbon atoms is 0.05 to 0.20 as measured by surface analysis using X-ray photoelectron spectroscopy.

[2] The carbon catalyst according to [1] above, wherein the ratio of the number of nitrogen atoms to the number of carbon atoms is 0.005 to 0.30 as measured by the surface analysis.

[0014]

[3] An electrode for redox flow batteries, comprising the carbon catalyst according to [1] or [2] above.

[4] A redox flow battery comprising the electrode according to [3] above.

[5] The redox flow battery according to [4] above, being a vanadium redox flow battery.

Advantageous Effects of Invention

[0015]

The carbon catalyst for redox flow battery electrodes according to the present invention is able to facilitate electrode reactions in redox flow batteries and has excellent durability.

The use of the carbon catalyst for redox flow battery electrodes according to the present invention can therefore provide an electrode for redox flow batteries that exhibits good charge-discharge cycle characteristics and can also increase the energy density of redox flow batteries. The carbon catalyst for redox flow battery electrodes according to the present invention is, in particular, suitably applicable to vanadium redox flow batteries.

Brief Description of Drawing

[0016]

[Fig. 1] Fig. 1 schematically shows the outline of the configuration of an exemplary redox flow battery.

Description of Embodiments

[0017]

Hereinafter, the present invention will be described in detail.

[Carbon catalyst]

The carbon catalyst for redox flow battery electrodes according to the present invention is characterized in that a ratio of the number of oxygen atoms to the number of carbon atoms (O/C ratio) is 0.05 to 0.20 as measured by surface analysis using X-ray photoelectron spectroscopy (XPS).

The carbon catalyst having such surface characteristics is able to facilitate electrode reactions in redox flow batteries, improve the charge-discharge characteristics, and provide satisfactory durability.

Specifically, the surface analysis using XPS can be conducted by the method described in Examples below.

[0018]

The carbon catalyst according to the present invention is applicable both to redox flow batteries employing aqueous electrolyte solutions and to those employing non-aqueous electrolyte solutions. In view of the fact that the carbon catalyst has hydrophilicity, the catalyst is preferably applied to redox flow batteries employing aqueous electrolyte solutions.

Electrolyte solutions of redox flow batteries contain metal ions whose valence is variable, and examples include iron-chromium electrolyte solutions, iron-titanium electrolyte solutions, titanium-manganese electrolyte solutions, manganese-chromium electrolyte solutions, chromium electrolyte solutions, and

vanadium electrolyte solutions. Among these, vanadium electrolyte solutions ($\text{VO}^{2+}/\text{VO}_2^+$ ($\text{V}^{4+}/\text{V}^{5+}$), $\text{V}^{2+}/\text{V}^{3+}$) are preferred to obtain a high electromotive force. These electrolyte solutions are preferably acidic aqueous solutions such as aqueous solutions of sulfuric acid, hydrochloric acid, nitric acid, or phosphoric acid and more preferably aqueous sulfuric acid solutions, in order to be solutions in which the metal ions or oxometal ions as active materials are stably present.

Thus, the electrolyte solutions in the present invention are preferably vanadium electrolyte solutions; that is, it is preferable that the catholyte be an aqueous sulfuric acid solution containing $\text{VO}^{2+}/\text{VO}_2^+$ ($\text{V}^{4+}/\text{V}^{5+}$) and the anolyte be an aqueous sulfuric acid solution containing $\text{V}^{2+}/\text{V}^{3+}$.

[0019]

In the carbon catalyst, the O/C ratio as measured by the XPS surface analysis is 0.05 to 0.20, preferably 0.06 to 0.18, and more preferably 0.10 to 0.18. When the O/C ratio is within this range, the catalytic activity can be improved as a result of improvement in the wettability with electrolyte solutions, so that electrode reactions can be facilitated.

If the O/C ratio is less than 0.05, the wettability with electrolyte solutions is low, so that the energy density (current density) tends to decrease. If the O/C ratio is more than 0.20, the carbon is likely to be oxidized, which leads to a low durability.

[0020]

The carbon catalyst preferably has a ratio of the number of nitrogen atoms to the number of carbon atoms (N/C ratio) of 0.005 to 0.30, more preferably 0.010 to 0.10, as measured by the XPS surface analysis. When the N/C ratio is within this range, the catalytic activity of the carbon catalyst can be further improved without decrease in its electrical conductivity.

[0021]

The carbon catalyst is preferably such that at least one of iron and cobalt is detected as a metal component in the XPS surface analysis. The carbon catalyst containing such a metal component can exhibit excellent catalytic activity. From the viewpoint of obtaining high catalytic activity, the ratio of the total number of iron atoms and cobalt atoms to the number of carbon atoms is preferably 0.0001 to 0.010 and more preferably 0.0002 to 0.008.

[0022]

[Method for producing carbon catalyst]

The method for producing the carbon catalyst of the present invention as described above is not particularly limited. The carbon catalyst is preferably produced by a method in which a raw material containing a nitrogen atom containing organic substance and a metal is carbonized. An example of the method is a carbonization method consisting successively of mixing of a transition metal powder into a solution of a nitrogen atom-containing resin, drying, infusibilizing treatment, baking, acid washing, and re-baking. Specifically, the carbon catalyst can be produced by the method as described in Examples below.

[0023]

The organic substance in the raw material subjected to carbonization is not particularly limited as long as it is a nitrogen-atom containing organic substance that can be carbonized, and any one or more of such substances can be used. For example, at least one selected from organic polymer compounds such as thermosetting resins and thermoplastic resins and low molecular weight organic compounds can be used. Biomass can also be used.

Specific examples of the organic substance include: cyclic organic compounds such as pyrrole, imidazole, pyridine, piperidine, triazole, and tetrazole; derivatives of these cyclic organic compounds; and polymer compounds such as polyacrylonitrile (PAN), acrylonitrile-methacrylic acid copolymer (PAN-PMA), polyimide, urea resin, polyaniline, and polypyrrole. Among these, acrylonitrile-methacrylic acid copolymer (PAN-PMA) and polyacrylonitrile (PAN) are preferably used.

[0024]

A transition metal is preferably contained as the metal in the raw material subjected to carbonization, from the viewpoint of effective improvement in catalytic activity of the carbon catalyst. Among transition metals, scandium, titanium, vanadium, chromium, manganese, iron, cobalt, nickel, zinc, or copper is preferred, iron, cobalt, or copper is more preferred, and iron or cobalt is even more preferred. These may be contained alone, or two or more thereof may be contained.

The metal may be added by itself or in the form of a metal compound. Examples of the metal compound include a metal salt, a metal oxide, a metal hydroxide, a metal nitride, a metal sulfide, a metal carbide, and a metal complex.

[0025]

The content of the metal in the raw material subjected to carbonization is not particularly limited as long as the carbon catalyst can exhibit desired catalytic activity. The content of the metal is preferably 1 to 90 parts by mass and more preferably 2 to 80 parts by mass relative to 100 parts by mass of the organic substance.

[0026]

The raw material may further contain other organic substances, other metals, and other components to the extent that the catalytic activity of the carbon catalyst is not impaired. Examples include: a curing agent and a crosslinking agent for the organic substance; a catalyst for the curing reaction or crosslinking reaction; and an electrically conductive carbon material such as carbon black or graphite powder.

[0027]

The raw material is carbonized by heating to or above the temperature at which the carbonization can be accomplished. The heating temperature for carbonization may be set as appropriate depending on, for example, the type and composition of the raw material used and can be, for example, 300 to 3000°C. The raw material is preferably heated to the temperature within 700 to 2000°C. The carbonization is preferably performed in an atmosphere of inert gas such as nitrogen.

The rate of temperature increase to the carbonization temperature and the duration during which the material is held at the carbonization temperature are not particularly limited as long as the carbonization can proceed sufficiently. For example, the rate of temperature increase can be 0.5 to 300°C/min, and the duration can be 5 minutes to 24 hours.

[0028]

Depending on, for example, the mode in which the carbon catalyst is to be used, the carbon catalyst obtained by carbonization of the raw material may be subjected to a further heat treatment or a process such as washing or crashing to further improve the catalytic activity.

The washing is a process for decreasing the metal content in the carbon catalyst and, for example, washing with an acid is preferred. The acid used in the washing is not particularly limited as long as the effect of decreasing the metal content can be achieved. Examples of the acid include concentrated hydrochloric acid, concentrated nitric acid, and concentrated sulfuric acid. The

method for the washing is not particularly limited. For example, the washing can be accomplished by the method of keeping the carbon catalyst immersed in an acid-containing solution.

To obtain the same effect, the washing may be replaced by electrolysis.
[0029]

The washing is preferably followed by a heat treatment in which the carbon catalyst is held at a predetermined temperature. The heat treatment temperature is preferably 300 to 3000°C and more preferably 700 to 2000°C. The rate of temperature increase to the heat treatment temperature can be, for example, 0.5 to 300°C/min. The duration during which the carbon catalyst is held at the heat treatment temperature can be, for example, 5 minutes to 24 hours. This heat treatment is preferably performed in an atmosphere of inert gas such as nitrogen gas.

[0030]

The crushing can be accomplished by a commonly known crushing means such as a ball mill or bead mill and can yield the carbon catalyst in the form of particles or powder.

[0031]

The carbon catalyst obtained as above may further be surface-treated with a 0.01 to 13 mol/L aqueous sulfuric acid solution at 35 to 90°C. In this surface treatment, the carbon catalyst is immersed in the aqueous sulfuric acid solution.

Even when the O/C ratio of the obtained carbon catalyst as measured by the XPS surface analysis is low, the O/C ratio can be increased by the above surface treatment using sulfuric acid to achieve improvement in durability and facilitation of electrode reactions. Furthermore, this surface treatment can serve to further decrease the metal component that remains in the carbon catalyst without being sufficiently removed even by the washing described above.

The use of nitric acid instead of sulfuric acid, although providing an increase in the O/C ratio, leads also to an increase in the N/C ratio as measured by the XPS surface analysis and a decrease in durability.

[0032]

In the surface treatment, the treatment temperature is more preferably 45 to 80°C and the concentration of the aqueous sulfuric acid solution is more preferably 1 to 4 mol/L, from the viewpoint of effective increase in the O/C ratio.

The surface treatment time may be set as appropriate depending on the

treatment temperature and the concentration of the aqueous sulfuric acid solution. From the viewpoint of securing sufficient time to obtain a satisfactory effect on increase in the O/C ratio, the surface treatment time is preferably 1 to 28 days.

[0033]

[Form]

The form of the carbon catalyst of the present invention is not particularly limited. It is preferable that crushing be performed in a process step subsequent to carbonization to obtain a powdered or particulate carbon catalyst.

Alternatively, a fibrous carbon catalyst may be obtained by processing the raw material into fibers using a commonly known spinning method and then by carbonizing the fibers. A sheet-shaped carbon catalyst can also be formed by making a woven fabric or the like using the fibrous carbon catalyst.

[0034]

[Electrode]

An electrode for redox flow batteries according to the present invention comprises the above carbon catalyst of the present invention.

The use of the carbon catalyst allows to form an electrode in which the electrode reaction is facilitated and which exhibits good charge-discharge cycle characteristics. Additionally, the improvement in electrode characteristics permits the electrode to be thinner than conventional electrodes.

[0035]

The electrode material to which the carbon catalyst of the present invention is applied is not particularly limited and may be a commonly known electrode for redox flow batteries. Typically, a carbon electrode is used. Among carbon electrodes, glassy carbon, carbon felt, and carbon fiber non-woven fabric are suitably used, since an electrode having a large surface area and many electrode reaction fields is preferred.

[0036]

The method for producing an electrode comprising the carbon catalyst is not particularly limited. It is preferable to fix the carbon catalyst to the surface of the electrode material with a binder. Examples of methods that can be used include: a method in which a liquid containing the carbon catalyst and the binder is applied to the surface of the electrode material with a coater or spray; a method in which the electrode material is immersed in a liquid containing the

carbon catalyst and the binder and then dried to fix the binder to the surface of the electrode material. With these methods, the carbon catalyst can be fixed uniformly over the entire surface of the electrode material.

The amount of the carbon catalyst fixed to the surface of the electrode material is not particularly limited as long as the amount is sufficient to obtain a satisfactory catalytic activity. The amount may be set as appropriate depending on, for example, the form of the electrode.

[0037]

[Redox flow battery]

A redox flow battery of the present invention comprises the above electrode. The electrode may be used either as the cathode of the redox flow battery or as the anode of the redox flow battery.

The use of the electrode comprising the carbon catalyst of the present invention allows obtaining a redox flow battery in which the electrode reaction is facilitated and which has a reduced cell resistance. Additionally, the facilitation of the electrode reaction can be achieved without the use of an expensive catalyst based on a noble metal such as platinum, and a redox flow battery having higher practical use can be provided at a low cost.

The carbon catalyst of the present invention is, in particular, suitably applicable to vanadium redox flow batteries.

Examples

[0038]

Hereinafter, the present invention will be described in more detail by Examples. The present invention is not limited by these Examples.

[0039]

[Production of carbon catalyst]

(Example 1)

A four-necked flask was charged with 30.93 g of acrylonitrile (available from Wako Pure Chemical Industries, Ltd.), 4.07 g of methacrylic acid (available from Wako Pure Chemical Industries, Ltd.), and 300 mL of pure water, and the contents of the flask were bubbled with nitrogen gas for 15 minutes. This flask was heated in an oil bath at 70°C, and a solution of 100 mg of potassium peroxodisulfate (available from Wako Pure Chemical Industries, Ltd.) dissolved in 50 ml of pure water was added into the flask, the contents of which were

stirred in a nitrogen gas atmosphere to allow polymerization to proceed for 4 hours. After leaving to cool, the resulting milky white solution was concentrated, and the concentrate was vacuum-dried at 60°C to obtain about 20 g of polyacrylonitrile-polymethacrylic acid copolymer (PAN-PMA).

[0040]

1.0 g of PAN-PMA obtained above was dissolved in 15 g of N,N-dimethylformamide to prepare a solution A. Additionally, 1.0 g of 2-methylimidazole and 5.78 g of zinc chloride were dissolved in 15 g of N,N-dimethylformamide to prepare a solution B. The solutions A and B were then mixed, and 0.187 g of iron powder was further added and mixed with the solution mixture.

[0041]

The resulting liquid mixture was vacuum-dried at 60°C for a whole day. The mixture resulting from the vacuum-drying was heated in air, and the temperature of the mixture was increased from room temperature (25°C) to 150°C over 30 minutes and then continuously increased from 150°C to 220°C over 2 hours. After that, the mixture was held at 220°C for 3 hours to infusibilize the mixture and thus prepare a carbonization raw material.

After leaving to cool, the carbonization raw material prepared was crushed with a planetary ball mill (P-7, available from Fritsch Japan Co., Ltd.).

[0042]

The crushed carbonization raw material was placed in a quartz tube and carbonized by heating the material in an image furnace in a nitrogen gas atmosphere, increasing the temperature of the crushed raw material from room temperature (25°C) to 1100°C at a temperature increase rate of 50°C/min and holding the material at 1100°C for 1 hour.

After leaving to cool, the resulting carbonized product was crushed with the same planetary ball mill as used above and with a bead mill (RMB-type batch bead mill, available from AIMEX Co., Ltd.).

[0043]

20 mL of concentrated hydrochloric acid was added to 1.0 g of the crushed carbonized product, the mixture was stirred for 30 minutes to precipitate the carbonized product, and the supernatant was removed. In this way, the carbonized product was acid-washed. This was followed by washing with a 1:1 mixed solution of concentrated hydrochloric acid and distilled water, filtration,

further washing with distilled water, and filtration. After the filtrate was confirmed to be neutral with pH paper, the filtration residue was vacuum-dried. [0044]

The carbonized product obtained as above was placed in a quartz tube and heat-treated by heating the product with an image furnace in a nitrogen gas atmosphere, increasing the temperature of the product from room temperature (25°C) to 700°C at a temperature increase rate of 50°C/min and holding the product at 700°C for 1 hour.

After natural cooling, the carbonized product was obtained as a powdered carbon catalyst.

[0045]

(Example 2)

0.5 g of the carbon catalyst produced in Example 1, 20 mL of a 4 mol/L aqueous sulfuric acid solution, and a stirrer bar were placed in a 100-mL eggplant flask equipped with a cooling tube, and the contents of the flask were heated under reflux using an oil bath at 85°C for 14 days while the contents were stirred. This was followed by natural cooling and then filtration, after which cycles of washing with distilled water and filtration were repeated. After the filtrate was confirmed to be neutral with pH paper, the filtration residue was dried with a reduced-pressure dryer (5000 Pa) at 60°C overnight.

The dried filtration residue was crushed with a mortar to a particle size of about 1 μm or less, giving a surface-treated carbon catalyst.

[0046]

(Examples 3 to 5)

Surface-treated carbon catalysts were obtained in the same manner as in Example 2, except that the temperature and time for heating under reflux were changed to those shown in Table 1 below.

[0047]

(Example 6)

0.5 g of the carbon catalyst produced in Example 1, 20 mL of a 4 mol/L aqueous sulfuric acid solution, and a stirrer bar were placed in a 20-mL sample vessel made of glass, and the contents of the vessel were stirred at room temperature (25°C) for 28 days. This was followed by filtration, after which the subsequent procedures were performed in the same manner as in Example 2, giving a surface-treated carbon catalyst.

[0048]

(Comparative Example 1)

0.5 g of the carbon catalyst produced in Example 1 was placed in a quartz tube and heat-treated by heating the catalyst with an image furnace in a nitrogen gas atmosphere, increasing the temperature of the catalyst from room temperature (25°C) to 1300°C at a temperature increase rate of 50°C/min and holding the catalyst at 1300°C for 1 hour.

After natural cooling, the resulting heat-treated product was crushed with a mortar to a particle size of about 1 μm or less, giving a surface-treated carbon catalyst.

[0049]

(Comparative Example 2)

A surface-treated carbon catalyst was obtained in the same manner as in Example 2, except that a 16 mol/L aqueous nitric acid solution was used instead of the 4 mol/L aqueous sulfuric acid solution and that the time for heating under reflux was 1 day.

[0050]

[Measurements and evaluations]

The carbon catalysts produced in Examples and Comparative Examples above were subjected to the following measurements and evaluations.

[0051]

(O/C ratio and N/C ratio)

The O/C ratio and N/C ratio in the surface of the carbon catalyst were determined by XPS surface analysis. The XPS surface analysis was conducted with an X-ray photoelectron spectrometer (AXIS NOVA, available from Kratos Analytical Ltd.) using AlK α ray (10 mA, 15 kV) as the X-ray source. For core-level photoelectron spectra of the various atoms in the surface of the carbonaceous particles, the binding energy was corrected by assuming the peak of the C1s spectrum to be at 284.5 eV. The O/C ratio and N/C ratio in the surface of the carbonaceous particles were calculated from the peak area of each spectrum and the detection sensitivity factor. The lower quantification limit was 0.0001.

[0052]

(Durability)

The durability was evaluated by an acid-resistance test, in which 500 mg of

the carbon catalyst was added to 20 mL of a 4 mol/L aqueous sulfuric acid solution and the mixture was stirred at 45°C for 3 hours.

The suspension obtained after the test was filtered to collect the carbon catalyst, and the carbon catalyst was subjected to repeated cycles of washing with distilled water and filtration. After the filtrate was confirmed to be neutral with pH paper, the carbon catalyst was dried with a reduced-pressure dryer (5000 Pa) at 60°C overnight. Subsequently, the dried carbon catalyst was collected, and its mass was measured. A rating of "A" was given when the change ratio in mass from the mass of the carbon catalyst before test was 5% or less, while a rating of "B" was given when the change in mass was more than 5%. [0053]

(Oxidation-reduction potential measurement)

First, the carbon catalysts prepared in Examples and Comparative Examples above were used to prepare electrode samples by the method described below. Each of the prepared electrode samples was subjected to oxidation-reduction potential measurement by cyclic voltammetry under the conditions listed below, and thus the oxidation-reduction potential difference, oxidation current density, and reduction current density were determined. [0054]

<Preparation of electrode sample>

500 μ L of a water-isopropyl alcohol mixture (at a mass ratio of 8:2) and 50 μ L of a 5 mass% aqueous solution of Nafion (registered trademark) were added to 5.0 mg of the carbon catalyst, and the catalyst was ultrasonically dispersed with an ultrasonication bath for 10 minutes. The dispersion was then stirred and mixed with a homogenizer (Sonifier (registered trademark) MODEL S-150D, available from BRANSON) for 2 minutes to prepare a catalyst slurry.

A drop of 1.382 μ L of this catalyst slurry was placed on a glassy carbon rotating disk electrode (diameter: 4 mm, length: 2.6 cm) using a microsyringe, and the slurry was spread over the entire surface of the electrode with the tip of the microsyringe. The slurry was then dried with hot air from a dryer to fix the carbon catalyst to the surface of the electrode at a density of 0.1 mg/cm².

A drop of a 0.5 mol/L aqueous sulfuric acid solution was placed on this electrode with the fixed carbon catalyst, and degassing was performed to replace the air in the electrode with the 0.5 mol/L aqueous sulfuric acid solution. In this way, the electrode sample was prepared.

[0055]

<Measurement conditions>

Working electrode: Rotating electrode employing the electrode sample prepared as above

Counter electrode: Platinum electrode

Reference electrode: Reversible hydrogen electrode (0.5 mol/L aqueous sulfuric acid solution)

Electrolyte solution: Aqueous vanadium-sulfuric acid solution (60 mL) containing 0.1 mol/L of vanadium ions (V^{4+}) and 2 mol/L of sulfate ions (SO_4^{2-})

Temperature: Room temperature (25°C)

Atmosphere: Degassed with nitrogen gas at 0.35 L/min for 10 minutes and then flowed.

Sweep potential: 1.5 to 0.3 V (versus reference electrode)

Sweep speed: 0.05 V/s

Sweep cycles: 5 cycles

[0056]

A smaller oxidation-reduction potential difference as determined by the above cyclic voltammetry indicates that the reversibility of the oxidation-reduction reaction is higher. Larger absolute values of the oxidation current density and reduction current density indicate that the catalytic activity is higher.

[0057]

The results of the measurements and evaluations described above are collectively shown in Table 1 below.

In Comparative Example 2, the oxidation-reduction peak was difficult to measure by the cyclic voltammetry for oxidation-reduction potential measurement, and the oxidation-reduction potential difference, oxidation current density, and reduction current density were not able to be determined.

[0058]

Table 1

	O/C ratio	N/C ratio	Durability	Oxidation-reduction potential difference (mV)	Oxidation current density (mA/cm ²)	Reduction current density (mA/cm ²)	Surface treatment
Example 1	0.07	0.016	A	181	4.548	-4.987	
Example 2	0.15	0.015	A	150	7.729	-6.200	85°C, 4 mol/L sulfuric acid, 14 days
Example 3	0.18	0.013	A	158	8.631	-6.701	85°C, 4 mol/L sulfuric acid, 28 days
Example 4	0.09	0.016	A	166	5.467	-4.905	85°C, 4 mol/L sulfuric acid, 1 day
Example 5	0.12	0.018	A	152	6.278	-5.543	45°C, 4 mol/L sulfuric acid, 14 days
Example 6	0.17	0.014	A	155	4.695	-5.009	25°C, 4 mol/L sulfuric acid, 28 days
Comparative Example 1	0.02	0.006	A	347	2.162	-2.910	1300°C, heat treatment
Comparative Example 2	0.25	0.032	B	-	-	-	85°C, 16 mol/L nitric acid, 1 day

[0059]

As seen from the results shown in Table 1, the oxidation-reduction potential difference was smaller when the O/C ratio was 0.05 to 0.20 (Examples 1 to 6) than when the O/C ratio was less than 0.05 (Comparative Example 1), which leads to the conclusion that the reversibility of oxidation-reduction reaction is higher when the O/C ratio is 0.05 to 0.20. Additionally, the absolute values of the oxidation current density and reduction current density were larger in Examples 1 to 6, which leads to the conclusion that the catalytic activity is higher in Examples 1 to 6.

Furthermore, comparison of Example 1 with Examples 2 to 6 leads to the conclusion that a predetermined surface treatment with sulfuric acid can provide further improvements in the reversibility of oxidation-reduction reaction and the catalytic activity.

Reference Signs List

[0060]

- 1 Cell
- 1a Cathode cell
- 1b Anode cell
- 2 Catholyte tank
- 3 Anolyte tank
- 4 Separation membrane

- 5 Cathode
- 6 Anode
- 7 Catholyte pump
- 8 Anolyte pump

CLAIMS:

1. A carbon catalyst for redox flow battery electrodes, wherein the carbon catalyst has a ratio of the number of oxygen atoms to the number of carbon atoms of 0.05 to 0.20, and a ratio of the number of nitrogen atoms to the number of carbon atoms of 0.005 to 0.30, in each case as measured by surface analysis using X-ray photoelectron spectroscopy.
2. An electrode for redox flow batteries, comprising the carbon catalyst according to claim 1.
3. A redox flow battery comprising the electrode according to claim 2.
4. The redox flow battery according to claim 3, being a vanadium redox flow battery.

[Fig. 1]

