METHOD FOR TREATING SCR CATALYSTS HAVING ACCUMULATED IRON COMPONDS

Abstract: The present disclosure relates to methods for treating deactivated SCR catalysts having an increased SO2/SO3 conversion rate as a result of the accumulation of one or more iron compounds. The methods are characterized in that the catalysts are treated with an aqueous solution of an acidic reactive salt or hydrogen fluoride, with the addition of at least one antioxidant.
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TECHNICAL FIELD OF THE INVENTION

The present disclosure describes novel methods of treating and regenerating selective catalytic reduction (SCR) catalysts having accumulations of one or more iron compounds.

BACKGROUND OF THE INVENTION

High temperature thermal processes, for example, generation of steam for the production of electricity in power plants utilizing fossil fuels, often create environmentally harmful by-products. These compounds, including nitrous oxides (NOx) and sulfur dioxide (SO2), must be removed from the flue gases of the high temperature thermal processes before the gases are discharged into the environment, for example before exiting the power plant and contacting the environment.

Desulfurization of the flue gas, for example removal of SO2, may be carried out by applying known methods in which the SO2 produced in the combustion process is oxidized to SO3. This is done prior to exposure of the flue gases to the reduction catalyst. The SO3 may then be absorbed into an alkaline solution and removed from the process, usually in the form of gypsum.

The standard for removing nitrous oxides from flue gases is the selective catalytic reduction (SCR) process, where a reducing agent, typically ammonia, is injected and mixed into the flue gases, and sent through a catalytic reaction chamber, where the catalyst facilitates the reduction of NOx by the reducing agent to form elemental nitrogen (N2) and water.

One undesired side reaction between the SCR catalyst and the constituents of the flue gas is the conversion of remaining SO2 to SO3. For example, the SO2 remaining in the flue gas may be partially oxidized to SO3, which may then react with water in the flue gas to produce sulfuric acid (H2SO4). Sulfuric acid in the flue gas stream may result in corrosion of steel surfaces, for example surfaces below the dew point of the sulfuric acid, in equipment downstream of the catalyst. In addition, emission of H2SO4 aerosol particles into the atmosphere may also be undesired for environmental reasons.

The catalysts, which in certain embodiments are referred to as DeNOx catalysts, may be constructed of titanium dioxide containing the oxides of transition metals, such as, for example, vanadium, molybdenum, and tungsten, to act as catalytically active components. In specific embodiments, the catalysts may be arranged on plates, in honeycomb fashion, or as a...
corrugated structure and are generally placed parallel to the direction of the flow of the flue gas. However, during operation of the power plant, the catalyst undergoes a loss of activity and efficiency, for example, due to plugging with fly ash and/or deactivation of the active components by certain compounds contained in the flue gas. Among these compounds are, for example, sodium (Na), potassium (K), arsenic (As), and phosphorous (P) based salts, as well as other compounds.

In addition to these compounds, iron compounds and/or iron salts may physically and chemically bond to the catalyst surface further reducing the performance of the catalyst. Research suggests that iron ions may be the main cause of the enhanced and undesired conversion of SO₂ to SO₃ during the regular operation of the SCR catalyst in the power plant. Iron contamination may come from a variety of sources, including the fuel burned in the power plant. For example, depending on the origin and age of coal, the natural iron content may range from about 1% to about 8% by weight, relative to the total amount of the mineral components in the coal.

It is generally known that during the regeneration of SCR catalysts, inorganic acids, such as sulfuric acid (H₂SO₄) and hydrochloric acid (HCl), may be used to clean and restore the catalyst, such as by soaking a step and a neutralizing step. Inorganic acids are typically odorless, which is another advantage to their use. Sulfuric acid, in particular, is relatively inexpensive and commercially available. In addition, sulfuric acid is also used because SO₂ and SO₃ may also be present in the flue gas and collected by the catalyst during operation in the power plant facility and when a catalyst is submerged in water or another aqueous solution, the SO₂ or SO₃ is removed from the catalyst and forms a diluted sulfuric acid solution. However, treating a catalyst with sulfuric acid has disadvantages since the H₂SO₄ in solution. However, treating a catalyst with sulfuric acid has disadvantages since the H₂SO₄ in a diluted aqueous solution also corrodes the steel casings of the catalyst. Corrosion of the catalyst casings may also result in release of water soluble iron compounds that can penetrate the pores of the SCR catalyst, further enhancing the undesired SO₂ to SO₃ conversion process.

The removal of iron contaminants from a DeNOx catalyst has been described in U.S. Patent No. 7,569,506 in which the catalyst is placed in a reaction solution comprising an aqueous solution of an inorganic or organic acid with the addition of one or more antioxidants. Inorganic acids, namely hydrochloric acid, phosphoric acid, nitric acid, and, in particular, sulfuric acid, are described. Organic acids, such as relatively strong organic acids, including oxalic acid, citric acid, malonic acid, formic acid, chloroacetic acid, and benzoic acid, sulfonic acids were also used. Although the methods described in this reference were effective in removing iron accumulation on the catalyst, the strong acids described in the reference also in removing iron accumulation on the catalyst, the strong acids described in the reference also
liberated iron ions from steel substrates and the steel casings of the catalyst. These iron ions can then penetrate the pores of the catalyst, potentially enhancing the undesired SO₂ to SO₃ conversion. Thus, there is a need for alternative methods for regeneration of SCR catalysts to remove or minimize contamination by iron compounds and provide optimum DeNOx performance of the catalyst while minimizing or reducing the SO₂ to SO₃ conversion process within the flue gas stream.

BRIEF DESCRIPTION OF THE FIGURES

The following figures assist in clarifying certain embodiments of the invention but are not intended to limit the invention to the examples described in the figures. Wherein Figures 1a, 1b, and 1c illustrate variability of pH as a function of the concentration of ammonium salt and ascorbic acid.

SUMMARY OF THE INVENTION

The present disclosure provides methods of regenerating an SCR catalyst, such as a DeNOx catalyst, having an accumulation of one or more iron compounds. The method comprises treating an SCR catalyst having an accumulation of one or more iron compounds with at least one aqueous solution comprising an acidic reactive component, such as ammonium salt or hydrogen fluoride, and at least one antioxidant, and removing at least a portion of the one or more iron compounds accumulated on the SCR catalyst.

In another embodiment, the present disclosure provides a method of regenerating a DeNOx catalyst having an accumulation of one or more iron compounds. The method comprises treating a DeNOx catalyst having an accumulation of one or more iron compounds with a substantially aqueous solution having a pH ranging from about 2.0 to about 5.0 and comprising from about 0.1% to about 4.0% by weight of an acidic reactive ammonium salt or about 0.5% to about 2.0% by weight of ascorbic acid or derivatives thereof, and removing at least a portion of the one or more iron compounds accumulated on the DeNOx catalyst, washing the treated DeNOx catalyst with water or an aqueous solution, and drying the DeNOx catalyst. Still other embodiments of the present disclosure provide for a regenerated SCR catalyst that has been regenerated according to the methods described herein.
DETAILED DESCRIPTION OF THE INVENTION

The present disclosure provides for methods of regenerating an SCR catalyst that has become contaminated with one or more iron compounds. SCR catalysts may become contaminated with iron compounds, for example, through the burning of fossil fuels that have a high concentration of iron compounds or through the leaching of iron compounds from power plant components. These iron based compounds may be a main contributor to the increase of the SO₂/ SO₃ conversion rate during the catalyst usage cycle. Iron compounds can both SO₂/ SO₃ conversion rate during the catalyst usage cycle. Iron compounds can both

become diminished catalytic performance and increase in SO₂/ SO₃ conversion rate. For example, iron compounds deposited on the catalyst can chemically react within the catalyst may result in the compounds deposited on the catalyst of chemically reacted with the catalyst may result in the oxidation of SO₂ in the flue gas to SO₃ which can then react with water in the flue gas to form sulfuric acid. The sulfuric acid may then react with downstream components of the flue gas emission system to corrode or oxidize those components, potentially leading to component failure.

Therefore, one goal of the present regeneration methods is to not only remove various contaminants responsible for the loss of catalytic activity but also maintain or decrease the SO₂/ SO₃ conversion rate potentially by removing the one or more iron compounds from the surface and pores of the catalyst during the regeneration process.

Other than the operating examples, or where otherwise indicated, all numbers expressing quantities of ingredients, processing conditions and the like used in the specification and claims are to be understood as being modified in all instances by the term "about". Accordingly, unless indicated to the contrary, the numerical parameters set forth in the following specification and attached claims are approximations that may vary depending upon the desired properties sought to be obtained. At the very least, and not as an attempt to limit the application of the doctrine of equivalents to the scope of the claims, each numerical parameter should at least be construed in light of the number of reported significant digits and parameter should at least be construed in light of the number of reported significant digits and by applying ordinary rounding techniques.

Notwithstanding that the numerical ranges and parameters setting forth the broad scope of the disclosure are approximations, the numerical values set forth in the specific examples are reported as precisely as possible. Any numerical values, however, may contain certain errors, such as for example, equipment and/or operator error, necessarily resulting from the standard deviation found in their respective testing measurements,arily resulting from the standard deviation found in their respective testing measurements.
Also, it should be understood that any numerical range recited herein is intended to include all sub-ranges subsumed therein. For example, a range of \( 1 \) to \( 10 \) is intended to include all sub-ranges between (and including) the recited minimum value of \( 1 \) and the recited maximum value of \( 10 \), that is, having a minimum value equal to or greater than \( 1 \) and a maximum value of \( 10 \), that is, having a minimum value equal to or greater than \( 1 \) and a maximum value of \( 10 \). Any patent, publication, or other disclosure material, in whole or in part, recited herein is incorporated by reference herein but only to the extent that the incorporated material herein is incorporated by reference herein but only to the extent that the incorporated material does not conflict with existing definitions, statements, or other disclosure material set forth in this disclosure. As such, and to the extent necessary, the disclosure as explicitly set forth herein supersedes any conflicting material incorporated herein by reference. Any material, or portion thereof, that is said to be incorporated by reference herein but which conflicts with existing definitions, statements, or other disclosure material set forth herein will only be incorporated to the extent that no conflict arises between that incorporated material and the existing disclosure material.

The present disclosure describes several different features and aspects of the invention with reference to various exemplary non-limiting embodiments. It is understood, however, that the invention embraces numerous alternative embodiments, which may be accomplished by combining any of the different features, aspects, and embodiments described herein, in any combination, that are one of ordinary skill in the art would find useful.

The present disclosure provides methods for treating SCR catalysts, such as DeNOx catalysts, that have become contaminated with one or more iron containing compounds. As used herein, the term "SCR catalyst" means a selective catalytic reduction catalyst designed to reduce the value of nitrogen oxides in the presence of ammonia to give elemental nitrogen and water. As used herein, the term "iron containing compound" means an iron containing compound, such as, but not limited to, iron compounds formed with a ferrous \((\text{Fe}^2+)\) ion or iron compounds formed with a ferric \((\text{Fe}^3+)\) ion (including mixtures of iron containing compounds). As used herein, an SCR catalyst that has been contaminated with iron has an accumulation of one or more iron compounds may have an iron concentration, either deposited on the catalyst or chemically reacted with the catalyst, of greater than 0.3% by weight of \( \text{Fe}_2\text{O}_3 \) (as measured by X-ray fluorescence) to the weight of the total catalyst material, not including the metal substrate on which a catalyst (i.e., plate type catalysts) may
be incorporated. Iron concentration may be measured as Fe₂O₃ concentration in the catalyst using X-ray fluorescent technology, which technology may also be used to measure concentrations of other components, including SiO₂, Al₂O₃, TiO₂, CaO, MgO, BaO, Na₂O, K₂O, SO₃, PO₄, V₂O₅, WO₃, MoO₃, As, and Cr₂O₃. Other methods may be used to measure iron concentration. In specific embodiments, the SCR catalyst may have an iron concentration ranging from about 0.3% to about 1% by weight of Fe₂O₃ in the catalyst as measured using X-ray fluorescence.

According to one embodiment, the present disclosure provides methods of regenerating an SCR catalyst, such as a DeNOx catalyst, wherein the catalyst has an accumulation of one or more iron compounds. The method may comprise treating the SCR catalyst having the accumulation of one or more iron compounds with a substantially aqueous solution comprising an acidic reactive salt or hydrogen fluoride, and at least one antioxidant; and removing at least a portion of the one or more iron compounds accumulated on the SCR catalyst. According to these embodiments, removing at least a portion of the one or more iron compounds provides a regenerated catalyst in which the rate of oxidation of SO₂ to SO₃ in flue gas is at least lower than the rate of oxidation of SO₂ to SO₃ in flue gas observed with a new catalyst.

In certain embodiments, the SCR catalyst may be treated with at least one acidic reactive salt and at least one antioxidant. As used herein, the term "acidic reactive salt" means any ionic salt that forms an acidic solution when dissolved in water (i.e., an aqueous solution having a pH of less than 7.0) and in which unsolvated water分子 are less than 35%. In specific embodiments, any ionic salt that forms an acidic aqueous solution having a pH of less than 5.0. Examples of acidic reactive salts suitable for use in various embodiments of the present disclosure include, but are not limited to ammonium-based acidic reactive salts, such as NH₄⁺, primary, secondary, or tertiary ammonium salts (primary, secondary or tertiary ammonium salts may be substituted with one or more alkyl group or other suitable organic group), having a suitable counterion, such as a counter anion selected from the group consisting of fluoride (F⁻), chloride (Cl⁻), bromide (Br⁻), iodide (I⁻), nitrate (NO₃⁻), dihydrogen phosphate (H₂PO₄⁻), hydrogen sulfate (HSO₄⁻), and sulfate (SO₄²⁻). In other embodiments, the acidic reactive salt may include salts of alkaline earth metals having a counterion selected from the group consisting of fluoride (F⁻), chloride (Cl⁻), bromide (Br⁻), iodide (I⁻), nitrate (NO₃⁻), dihydrogen phosphate (H₂PO₄⁻), hydrogen sulfate (HSO₄⁻), and sulfate (SO₄²⁻). As used herein, the term "substantially aqueous solution" means a solution that comprises water and at least one ionic compound.
(such as the acidic reactive salt) or hydrogen fluoride and at least one antioxidant compound (such as the acidic reactive salt) or hydrogen fluoride and at least one antioxidant compound dissolved in the water in such a concentration that the solution is still free flowing as a liquid (i.e., dissolved up to the saturation point). A substantially aqueous solution may further

5 comprise one or more other compounds such as an organic solvent, such as a polar organic solvent, for example an alcohol, provided the other compound(s) do not interfere with the ability (wettability) of the substantially aqueous solution to contact and wet the surface of the catalyst, including penetration of the solution into the pores of the catalyst. In certain embodiments, the concentration of the antioxidant may range from about 0.5% to about 2.0% by weight of the solution.

10 substantial, substantially aqueous solution, substantially aqueous solution comprising the acidic reactive salt in a concentration ranging from about 0.1% to about 10.0%. In other embodiments, the substantially aqueous solution may comprise the acidic reactive salt in a concentration ranging from about 0.1% to about 4.0% by weight, and in other embodiments from about 0.5% to about 2.0% by weight of the solution. In still other embodiments, the substantially aqueous solution may comprise sufficient acidic reactive salt to make the pH of the solution less than 5.0.

According to other embodiments, the SCR catalyst may be treated with a substantially aqueous solution comprising hydrogen fluoride and at least one antioxidant (i.e., a substantially aqueous hydrofluoric acid solution comprising at least one antioxidant). As used herein, the phrases "a substantially aqueous solution comprising hydrogen fluoride" and "a hydrofluoric acid solution" mean the same and may be used interchangeably. For example, according to these embodiments, the substantially aqueous solution may comprise hydrogen fluoride in a concentration ranging from about 0.05% to about 0.2% by weight of the solution. In a concentration ranging from about 0.05% to about 0.2% by weight of the solution.

In addition to the acidic reactive salt or hydrogen fluoride, the substantially aqueous solution further comprises at least one antioxidant. Suitable antioxidants may include, for example, enols or enediols such as for example ascorbic acid, carboxylic acids such as for example "ascorbic acid " or "ascorbic acid derivative thereof," for example citric acid and hydroxy carboxylic acids such as for example "tartaric acid," and salicylic acid or derivatives of any of these compounds. In specific embodiments, the antioxidant may be ascorbic acid or a derivative thereof. According to these embodiments, the substantially aqueous solution may contain the antioxidant, for example ascorbic acid or a derivative thereof, in a concentration of about 0.5% to about 10.0% by weight of the solution. In other embodiments, the concentration of the antioxidant may range from about 0.5% to about 2.0% by weight of the solution.
According to the embodiments where the substantially aqueous solution comprises either an acidic reactive salt or hydrogen fluoride, the acidic reactive salt or hydrogen fluoride is added in a concentration sufficient to bring the pH of the solution to less than 5.0, for example, in some embodiments the pH range may be from about 2.0 to about 3.0. While treating the catalyst with the substantially aqueous solution, the inventors believe that

Without intending to be bound by any theory of operation, the inventors believe that the antioxidant component may reduce and/or bind to the one or more iron compounds that the antioxidant compound may reduce and/or bind to the one or more iron compounds that have accumulated on the catalyst, thereby forming a more soluble form of the iron compound which is then solubilized in the aqueous solution and removed from the catalyst. For example, in those embodiments where the antioxidant comprises ascorbic acid, the ascorbic acid may reduce insoluble Fe^{2+} ions to Fe^{0} ions, which may then form a soluble metal/ligand complex with the ascorbic acid. In other embodiments, the antioxidant may reduce the iron compound and the iron compound may form a more soluble form of the iron compound with the antioxidant in the substantially aqueous solution. The antioxidant is believed to reduce and/or bind to the substantially aqueous solution. The antioxidant/iron complex of the reduced form of iron may then be solubilized and washed away from the surface of the catalyst. However, without intending to be bound by any theory, it is believed that binding, reduction and/or solubilization of the antioxidant/iron complex requires the presence of the acidic reactive salt or hydrogen fluoride in the substantially aqueous solution. For example, in the absence of the acidic reactive salt or hydrogen fluoride, the binding constant of the antioxidant to the iron compound may not be sufficient to form a stable, soluble complex compared to the binding constant of the iron to the catalyst. That is, the binding constant of the iron to the catalyst is higher than the binding constant of the antioxidant to the iron. In particular, it is believed that when the iron is chemically bound to the catalyst surface to a poison which itself is attached to the catalyst surface, a stronger agent, such as the acidic reactive salt or the hydrogen fluoride may be necessary to break any of such chemical, ionic or physical bonds or attractions, allowing the free iron compound to react with the antioxidant.

According to specific embodiments, treating and regenerating the SCR catalyst may be affected at any temperature suitable to remove the one or more iron compounds into the substantially aqueous solution. In one embodiment, treating the SCR catalyst with the substantially aqueous solution may be at a temperature ranging from about 0°C to about 100°C. In another embodiment, treating the SCR catalyst with the substantially aqueous solution may be at a temperature ranging from about 0°C to about 60°C. In another embodiment, treating the SCR catalyst with the substantially aqueous solution may be at a temperature ranging from about 20°C to about 60°C. In yet another embodiment, treating the SCR catalyst with the substantially aqueous solution may involve at least partially immersing the catalyst in a tank or basin holding the substantially aqueous solution.
In other embodiments, treating the catalyst with the substantially aqueous solution may include spraying the catalyst with the substantially aqueous solution. In still other embodiments, treating the catalyst with the substantially aqueous solution may flow the solution through one or more apertures of the catalyst, for example, but not limited to the process described in U.S. Published Application No. 2008/015800, the disclosure of which is incorporated in its entirety by this reference.

In particular embodiments, treating the SCR catalyst may comprise at least one of moving the catalyst and agitating the substantially aqueous solution, such as while the catalyst is at least partially immersed in a tank containing the substantially aqueous solution. For example, moving the catalyst and/or agitating the solution during the treating step may provide improved contact between the substantially aqueous solution and the surface of the catalyst and, therefore, with the one or more iron compounds on the surface of the catalyst or chemically bonded to the catalyst. In addition, moving the catalyst and/or agitating the solution containing the catalyst may further improve removal of the one or more iron compounds by helping to dislodge fly ash particulate matter and/or the iron compound (either by itself or after forming a complex with the antioxidant) from the surface of the catalyst.

According to certain embodiments, moving the catalyst may comprise at least one of raising and lowering the catalyst, moving the catalyst in one or more longitudinal, lateral, circular, or diagonal directions, pivoting the catalyst around at least one of a vertical, horizontal, or diagonal axis, or shaking or vibrating the catalyst. According to other embodiments, agitating the substantially aqueous solution may comprise at least one of stirring the solution, mixing the solution using mechanical mixers or air blowers, recirculating the solution, injecting the solution through one or more openings in the catalyst, spraying the solution on at least one surface of the catalyst, and treating the solution and the catalyst with ultrasonic or low frequency vibrations. In specific embodiments, treating the catalyst may include both moving frequency vibrations. In specific embodiments, treating the catalyst may include both moving the catalyst in the solution and agitating the solution.

According to other embodiments, after treating the SCR catalyst as described herein, the methods include removing at least a portion of the one or more iron compounds accumulated on the SCR catalyst. As described herein, the iron compounds may be removed after being treated with the substantially aqueous solution, such as by dissolving into the substantially aqueous solution or being dislodged into the substantially aqueous solution. As the catalyst is removed from the aqueous solution, the iron compounds and residues remain in the aqueous solution, and are therefore removed from the catalyst surface. According to certain aqueous embodiments, sufficient amounts of the iron compounds may be removed from the catalyst.
such that the activity of the catalyst (e.g., the SO$_2$/SO$_3$ conversion rate potential) is 
regenerated to the levels present in new catalysts, or lower, in specific embodiments. 
removing at least a portion of the one or more iron compounds results in a regenerated 
catalyst having less than 0.3% by weight of iron contamination, such as Fe$_3$O$_4$ contamination,
treating the SCR catalyst with the substantially aqueous solution as described herein.

Washing the SCR catalyst after the treating and removing steps may further remove residual iron or other contaminants on the catalyst surface, either by dissolving or dislodging the residual iron or other contaminants. The catalyst may be washed in this manner by exposing the catalyst surface to the substantially aqueous solution, or by moving the catalyst through the solution in a stirring or agitation manner. After washing, the catalyst may be dried, as such as by drying the catalyst in a drying oven at a temperature of at least 100°C. Alternatively, in those embodiments where the catalyst is not subjected to a washing step after the treating and removing steps, the catalyst may be dried directly after the step of removing at least a portion of the accumulated iron compounds, removing at least a portion of the accumulated iron compounds.

After treating the catalyst according to any of the various embodiments described in detail herein, the catalyst may be optionally further regenerated, for example, by impregnating the SCR catalyst with one or more catalytically active metal compounds. For example, during the intended use (i.e., removing one or more components from a power plant flue gas) and/or during various treating/regeneration processes of the catalyst, the catalytic activity of the catalyst may become diminished due to, for example, loss of activity of one or more catalytically active metal compounds in the SCR catalyst. Re-impregnating the regenerated SCR catalyst may include impregnating one or more catalytically active metal compounds selected from the group consisting of vanadium compounds, molybdenum compounds, and tungsten compounds, such as oxides of these metals. In other embodiments, the regenerated SCR catalyst may further be re-calcinated, for example, to improve the mechanical strength and structural integrity of the regenerated catalyst.

According to other embodiments, the present disclosure provides a method for regenerating a DeNOx catalyst having an accumulation of one or more iron compounds. The method according to these embodiments comprises treating the DeNOx catalyst having an accumulation of one or more iron compounds with a substantially aqueous solution having a pH ranging from about 2.0 to about 8.0 and comprising from about 0.1% to about 4.0% by weight of an acidic reactant ammonium salt, such as described herein, and from about 0.3% to about 2.0% by weight or ascorbic acid or a derivative thereof, removing at least a portion of the one or more iron compounds accumulated on the DeNOx catalyst; washing the treated DeNOx catalyst with water or other aqueous solution to remove any residual iron compounds of other contaminants; and drying the DeNOx catalyst. Specific embodiments of these steps are described in detail herein.
Other embodiments provide for a regenerated SCR catalyst that has been regenerated according to any of the methods set forth herein. The methods, described herein may take an SCR catalyst, such as a DeNOx catalyst, having an accumulation of one or more iron compounds, resulting in an increase in the conversion of SO2 to SO3 in flue gas, and regenerate the catalyst by removing at least a portion of the contaminating iron compounds, thereby decreasing the rate of conversion of SO2 to SO3 in flue gas. The regenerated catalyst, for example a regenerated DeNOx catalyst, that have been treated according to any catalyst, for example a regenerated DeNOx catalyst, that have been treated according to any one of the methods described herein may demonstrate the same or even lower oxidation of SO2 to SO3 compared to a new catalyst. For example, according to certain embodiments, an SO3 to SO2 compared to a new catalyst. For example, according to certain embodiments, an SCR catalyst, such as a DeNOx catalyst that has been treated according to an embodiment of SCR catalyst, such as a DeNOx catalyst that has been treated according to an embodiment of the present disclosure may display at least a 10% reduction of the SO2 to SO3 conversion rate compared to an untreated, iron contaminated catalyst.

The present invention set forth in the specification and defined by the claims will be better understood read in conjunction with the following non-limiting exemplary examples.

EXAMPLES

Example 1

In this example, a pre-test was performed to identify potential candidates for ammonium salts based on pH of the salt-ascorbic acid mixture. The pH of substantially aqueous solutions of various ammonium acids, reactive salts were examined. The salts examined in this example were: Ammonium Chloride (NH4Cl); Ammonium Fluoride (NH4F); Ammonium Nitrate (NH4NO3); and Ammonium Sulfate ((NH4)2SO4).

The salt concentrations were selected at 1%, 2% and 4% wt. and Ascorbic Acid concentration were 0.5%, 1%, and 2% wt. The effect on pH for aqueous solutions of these concentrations was 0.5%, 1%, and 2% wt. The effect on pH for aqueous solutions of these compositions are presented in Table 1. Figures 1a, 1b and 1c illustrate the effect on pH of various concentrations of ammonium salts and ascorbic acid at differing temperatures. Figure 1a shows that after an initial addition of 0.5% wt. Ascorbic Acid to a 1% salt solution the pH dropped from between 6 and 7 to between 2 and 3.5 for all salts and then remained almost constant until after adding up to 4% wt. Ascorbic Acid. Increasing the salt concentration further had no effect on the pH value. It can also be seen that NH4F had a significantly higher pH after Ascorbic Acid addition compared to the other salts. This effect increased with increasing NH4F concentration (compare Figs. 1a, 1d, and 1c). The lowest pH value for all ammonium fluoride solution was 4.40 for 1% NH4F and 2.0% Ascorbic Acid by weight (Fig. ammonium fluoride solution was 4.40 for 1% NH4F and 2.0% Ascorbic Acid by weight (Fig.
The lowest pH value an ammonium chloride solution was 2.78 for 4\% wt NH₄Cl and 2.0\% wt Ascorbic Acid (Fig. 1c). A similar value of pH 2.84 was achieved with the \( \text{NH}_4\text{Cl} \) and \( \text{NH}_2\text{NO}_3 \) solution containing 4\% \( \text{NH}_4\text{Cl} \) and 2.0\% wt Ascorbic Acid (Fig. 1c).

Although the corresponding \( (\text{NH}_4)_2\text{SO}_4 \) solution (4\% \( (\text{NH}_4)_2\text{SO}_4 \) and 2.0\% wt Ascorbic Acid) had a higher pH value of 3.55, it was decided to further investigate the acidic reactive ammonium salts NH₄Cl and \( (\text{NH}_4)_2\text{SO}_4 \).

Example 2

A Honeycomb SCR catalyst was tested for iron removal using combinations of Ammonium Chloride, Ammonium Sulfate and Hydrogen Fluoride with Ascorbic Acid. Approximately 30 g of the Honeycomb catalyst was placed in 270 mL of the solutions of interest (see Tables 3 through 2) and heated and mixed for 4 hrs in a water bath. Samples of interest were taken at the beginning of the test, after 1 hr and after 4 hrs and reserved for the solutions were taken at the beginning of the test, after 1 hr and after 4 hrs and reserved for pH, Fe\(^{2+}\) and Total Iron analysis using the analytical test methods set forth in Table 2.

As can be seen in Table 2, the range for Fe\(^{2+}\) analysis is very broad while the limit for Total Iron analysis is 3,000 mg/L. Dilutions of up to 1:200 were required in order to stay within the measurement range for Total Iron analysis. As a result, the measured Total Iron values were in some cases lower than the measured Fe\(^{2+}\) values. Total Iron is the sum of Fe\(^{2+}\) and Fe\(^{3+}\). Therefore, the Total Iron will be at least equal to or greater than the Fe\(^{2+}\) concentration. Certain experimental errors may be observed due to the imprecise nature of the visual comparison using Fe\(^{2+}\) teststrip method and the precise but highly diluted value observed using Total Iron measurement.

Results from measuring the pH, Fe\(^{2+}\) concentration and Total Iron concentration from SCR catalysts treated with the formulations in this Example are presented in Tables 3, 4, and 5, respectively. In general, the increase in Total Iron concentration was proportional to the increase in Ammonium Salt concentration. Increasing the Ascorbic Acid concentration from 0.3\% to 0.5\% by weight resulted in an increase of the Total Iron concentration after 4 hrs from 126 mg/L to 184 mg/L (47\%) for the 0.5\% wt \( (\text{NH}_4)_2\text{SO}_4 \) solution and from 122 mg/L to 199 mg/L (64\%) for the 0.5\% wt \( \text{NH}_4\text{Cl} \) solution. The Total Iron concentration in the test solution after 4 hrs was 257 mg/L with the combination of 0.2\% Hydrogen Fluoride and 0.5\% Ascorbic Acid by weight. A combination of 2\% \( (\text{NH}_4)_2\text{SO}_4 \) and 0.5\% Ascorbic Acid by weight showed a Total Iron value of 249 mg/L. A concentration of 230 mg/L of Total Iron weight showed a Total Iron concentration of 3,000 mg/L. A graphical representation of Total Iron concentration for various test conditions is presented in Figure 2.
Table 1: pH-Value of Ammonium Salts at Different Temperatures and Varying Ascorbic Acid Concentration

<table>
<thead>
<tr>
<th>Condition</th>
<th>1% wt NH₄F</th>
<th>1% wt NH₄Cl</th>
<th>1% wt (NH₄)₂SO₄</th>
<th>1% wt NH₄NO₃</th>
<th>2% wt NH₄F</th>
<th>2% wt NH₄Cl</th>
<th>2% wt (NH₄)₂SO₄</th>
<th>2% wt NH₄NO₃</th>
<th>4% wt NH₄F</th>
<th>4% wt NH₄Cl</th>
<th>4% wt (NH₄)₂SO₄</th>
<th>4% wt NH₄NO₃</th>
</tr>
</thead>
<tbody>
<tr>
<td>60°C</td>
<td>6.14</td>
<td>5.50</td>
<td>5.83</td>
<td>5.90</td>
<td>5.34</td>
<td>6.06</td>
<td>5.52</td>
<td>5.48</td>
<td>6.38</td>
<td>5.34</td>
<td>5.66</td>
<td>5.43</td>
</tr>
<tr>
<td>60°C + 0.5%wt Ascorbic Acid</td>
<td>4.68</td>
<td>2.79</td>
<td>3.32</td>
<td>2.92</td>
<td>5.08</td>
<td>3.10</td>
<td>3.51</td>
<td>3.01</td>
<td>5.64</td>
<td>3.07</td>
<td>3.65</td>
<td>3.01</td>
</tr>
<tr>
<td>60°C + 1.0%wt Ascorbic Acid</td>
<td>4.68</td>
<td>3.01</td>
<td>3.47</td>
<td>3.10</td>
<td>5.15</td>
<td>3.22</td>
<td>3.65</td>
<td>3.19</td>
<td>5.75</td>
<td>2.96</td>
<td>3.66</td>
<td>3.03</td>
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<tr>
<td>60°C + 2.0%wt Ascorbic Acid</td>
<td>4.40</td>
<td>2.87</td>
<td>3.23</td>
<td>2.85</td>
<td>5.13</td>
<td>3.04</td>
<td>3.46</td>
<td>2.91</td>
<td>5.60</td>
<td>2.78</td>
<td>3.55</td>
<td>2.84</td>
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</table>
Table 2: Overview of Analytical Methods

<table>
<thead>
<tr>
<th></th>
<th>Instrument</th>
<th>Method</th>
<th>Range/Limit</th>
</tr>
</thead>
<tbody>
<tr>
<td>pH</td>
<td>Accumet</td>
<td>Model AP85</td>
<td>-2.00-16.00pH/ Resolution 0.01pH</td>
</tr>
<tr>
<td>Fe$^{2+}$</td>
<td>MERCK</td>
<td>EM Quant 10004-1 Teststrip - visual comparison</td>
<td>0, 3, 10, 25, 50, 100, 250, 500 mg/L; 3mg/L</td>
</tr>
<tr>
<td>Iron, Total</td>
<td>HACH</td>
<td>FerroVer Method Colorimetric method – photometric reading</td>
<td>0 to 3,000mg/L / 0.03mg/L</td>
</tr>
</tbody>
</table>

Example 3 – Determination of Change in SO$_2$ to SO$_3$ Conversion Rate

In this Example, the change in the SO$_2$ to SO$_3$ conversion rate for an SCR catalyst treated according to one embodiment of the present disclosure was determined. The SO$_2$ to SO$_3$ conversion rate of the catalyst contaminated with iron was determined prior to and after treatment according to one embodiment of the present disclosure.

A sample log from a honeycomb SCR catalyst contaminated with iron was analyzed to determine the SO$_2$ to SO$_3$ conversion rate of the catalyst. The SO$_2$ to SO$_3$ conversion rate is determined according to the protocol recited in Chapter 4: SO$_2$ to SO$_3$ Oxidation of Electric Power, Research Institute, "Protocol for Laboratory Testing of SCR Catalyst: 2nd Edition," D. Broske, Technical Update, December 2007, pp 4-1 to 4-11. The SO$_2$ to SO$_3$ conversion rate of the untreated sample log was determined to be 0.752%. The sample log was then treated for 3 hours in an aqueous solution comprising 2% wt. (NH$_4$)$_2$SO$_4$ and 0.5% wt. Ascorbic Acid at a temperature from 40°C to 50°C. The sample log was then washed with deionized water for 1 hour in a cascading fashion. The Fe$^{2+}$ content of the treatment solution was measured to be 0.752%.

The sample log was then dried and calcined for 8 hours at 400°-425°C (750°-800°F) and the SO$_2$ to SO$_3$ conversion rate of the treated sample log was determined. The treated sample log displayed an SO$_2$ to SO$_3$ conversion rate of 0.201%, corresponding to a 73% reduction of the SO$_2$ to SO$_3$ conversion compared to the untreated iron contaminated log.

The untreated iron contaminated log had a SO$_2$ to SO$_3$ conversion rate of 0.752%, corresponding to a 73% reduction of the SO$_2$ to SO$_3$ conversion compared to the untreated iron contaminated log.
Table 3: pH-Value of Ammonium Salts at Different Temperatures and Varying Ascorbic Acid Concentration

<table>
<thead>
<tr>
<th></th>
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<th></th>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>Test Condition</td>
<td>HF 0.05% / AA 0.5%</td>
<td>HF 0.15% / AA 0.5%</td>
<td>HF 0.2% / AA 0.5%</td>
<td>NH₄Cl 0.5% / AA 0.5%</td>
<td>NH₄Cl 0.5% / AA 0.3%</td>
<td>NH₄Cl 0% / AA 0.5%</td>
<td>NH₄Cl 0% / AA 0.5%</td>
<td>(NH₄)₂SO₄ 0.5% / AA 0.3%</td>
<td>(NH₄)₂SO₄ 1% / AA 0.5%</td>
<td>(NH₄)₂SO₄ 2% / AA 0.5%</td>
<td></td>
<td></td>
</tr>
<tr>
<td>pH [-]</td>
<td>2.04</td>
<td>2.41</td>
<td>2.29</td>
<td>2.26</td>
<td>3.01</td>
<td>2.82</td>
<td>2.88</td>
<td>2.90</td>
<td>3.35</td>
<td>3.21</td>
<td>3.27</td>
<td>3.33</td>
</tr>
<tr>
<td>RT / 0hr</td>
<td>3.29</td>
<td>3.79</td>
<td>2.85</td>
<td>2.76</td>
<td>3.67</td>
<td>3.57</td>
<td>3.68</td>
<td>3.55</td>
<td>3.86</td>
<td>3.82</td>
<td>3.86</td>
<td>3.86</td>
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<tr>
<td>50 °C / 4hr</td>
<td>3.70</td>
<td>3.48</td>
<td>3.23</td>
<td>3.02</td>
<td>4.20</td>
<td>4.01</td>
<td>4.04</td>
<td>3.94</td>
<td>4.43</td>
<td>4.21</td>
<td>4.17</td>
<td>4.18</td>
</tr>
</tbody>
</table>

Table 4: Iron-Removal of Ammonium Salts at Different Temperatures and Varying Ascorbic Acid Concentration

<table>
<thead>
<tr>
<th></th>
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<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>Test Condition</td>
<td>HF 0.05% / AA 0.5%</td>
<td>HF 0.15% / AA 0.5%</td>
<td>HF 0.2% / AA 0.5%</td>
<td>NH₄Cl 0.5% / AA 0.3%</td>
<td>NH₄Cl 0.5% / AA 0.3%</td>
<td>NH₄Cl 0% / AA 0.5%</td>
<td>NH₄Cl 0% / AA 0.5%</td>
<td>(NH₄)₂SO₄ 0.5% / AA 0.3%</td>
<td>(NH₄)₂SO₄ 1% / AA 0.5%</td>
<td>(NH₄)₂SO₄ 2% / AA 0.5%</td>
<td></td>
<td></td>
</tr>
<tr>
<td>Fe²⁺ [mg/L]</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
</tr>
<tr>
<td>Fe³⁺ [mg/L]</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
</tr>
<tr>
<td>RT / 0hr</td>
<td>50</td>
<td>75</td>
<td>100</td>
<td>125</td>
<td>40</td>
<td>75</td>
<td>125</td>
<td>100</td>
<td>40</td>
<td>75</td>
<td>125</td>
<td>100</td>
</tr>
<tr>
<td>50 °C / 4hr</td>
<td>150</td>
<td>150</td>
<td>200</td>
<td>250</td>
<td>200</td>
<td>100</td>
<td>225</td>
<td>250</td>
<td>200</td>
<td>100</td>
<td>225</td>
<td>250</td>
</tr>
</tbody>
</table>
### Table 5: Total Iron-Removal of Ammonium Salts at Different Temperatures and Varying Ascorbic Acid Concentration

<table>
<thead>
<tr>
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<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>Test Condition</td>
<td>HF 0.05% / AA 0.5%</td>
<td>HF 0.1% / AA 0.5%</td>
<td>HF 0.15% / AA 0.5%</td>
<td>HF 0.2% / AA 0.5%</td>
<td>NH₄Cl 0.5% / AA 0.3%</td>
<td>NH₄Cl 0.5% / AA 0.5%</td>
<td>NH₄Cl 1% / AA 0.5%</td>
<td>NH₄Cl 2% / AA 0.5%</td>
<td>(NH₄)₂SO₄ 0.5% / AA 0.3%</td>
<td>(NH₄)₂SO₄ 0.5% / AA 0.5%</td>
<td>(NH₄)₂SO₄ 1% / AA 0.5%</td>
<td>(NH₄)₂SO₄ 2% / AA 0.5%</td>
</tr>
<tr>
<td>Iron Total [mg/L]</td>
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<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
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<td>0</td>
<td>0</td>
</tr>
<tr>
<td>Iron Total [mg/L]</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
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<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
<td>0</td>
</tr>
<tr>
<td>Iron Total [mg/L]</td>
<td>51</td>
<td>82</td>
<td>124</td>
<td>166</td>
<td>69</td>
<td>85</td>
<td>121</td>
<td>161</td>
<td>62</td>
<td>96</td>
<td>122</td>
<td>134</td>
</tr>
<tr>
<td>Iron Total [mg/L]</td>
<td>121</td>
<td>163</td>
<td>206</td>
<td>257</td>
<td>122</td>
<td>199</td>
<td>214</td>
<td>230</td>
<td>126</td>
<td>184</td>
<td>237</td>
<td>249</td>
</tr>
</tbody>
</table>

Table 6: Iron-Removal for Sample Log for SO₂ to SO₃ Conversion Rate Determination

<table>
<thead>
<tr>
<th></th>
<th>pH</th>
<th>Temperature [°C]</th>
<th>Fe²⁺ [mg/L]</th>
<th>SO₂→SO₃ conversion</th>
</tr>
</thead>
<tbody>
<tr>
<td>TreatmentStart</td>
<td>3.80</td>
<td>50</td>
<td>0</td>
<td>0.752%</td>
</tr>
<tr>
<td>TreatmentEnd</td>
<td>2.75</td>
<td>41</td>
<td>275</td>
<td></td>
</tr>
<tr>
<td>CascadeStart</td>
<td>5.50</td>
<td>53</td>
<td>3</td>
<td></td>
</tr>
<tr>
<td>CascadeEnd</td>
<td>5.40</td>
<td>48</td>
<td>15</td>
<td>0.201%</td>
</tr>
</tbody>
</table>
WE CLAIM

1. A method of regenerating an SCR catalyst having an accumulation of one or more iron compounds, the method comprising: treating an SCR catalyst having an accumulation of one or more iron compounds with a substantially aqueous solution comprising: an acidic reactive salt or hydrogen fluoride; and at least one antioxidant; and removing at least a portion of the one or more iron compounds accumulated on the SCR catalyst.

2. The method of claim 1, wherein the substantially aqueous solution comprises an acidic reactive salt and at least one antioxidant.

3. The method of claim 1 or 2, wherein the acidic reactive salt is an ammonium-based acidic reactive salt.

4. The method of any of claims 1 to 3, wherein the acidic reactive salt is an ammonium salt selected from the group consisting of ammonium chloride, ammonium fluoride, ammonium nitrate, ammonium sulfate, and mixtures of any thereof.

5. The method of any of claims 1 to 4, wherein the acidic reactive salt is ammonium chloride or ammonium sulfate.

6. The method of any of claims 1 to 5, wherein the concentration of the acidic reactive salt ranges from about 0.1 wt% to about 10.0 wt%, preferably about 0.5 wt% to about 2.0 wt%.

7. The method of any of claims 1 to 6, wherein the substantially aqueous solution comprises hydrogen fluoride in a concentration ranging from about 0.05 wt% to about 0.2 wt%.

8. The method of any of claims 1 to 7, wherein the substantially aqueous solution has a pH ranging from about 2.0 to about 5.0.
9. The method of any of claims 1 to 8, wherein the at least one antioxidant comprises ascorbic acid or derivatives thereof.

10. The method of claim 9, wherein the concentration of ascorbic acid or derivative thereof in the substantially aqueous solution ranges from about 0.5 wt% to about 10.0 wt%.

11. The method of any of claims 1 to 10, wherein the SCR catalyst is a DeNOx catalyst.

12. The method of any of claims 1 to 11, wherein treating the SCR catalyst comprises treating the SCR catalyst at a temperature ranging from about 0°C to about 100°C.

13. The method of any of claims 1 to 12, wherein treating the SCR catalyst further comprises at least one of moving the catalyst in the solution and agitating the substantially aqueous solution.

14. The method of claim 13, wherein moving the catalyst in the solution comprises at least one of raising and lowering the catalyst, pivoting the catalyst around a vertical, diagonal, or horizontal axis; moving the catalyst in a longitudinal direction; and shaking the catalyst.

15. The method of claim 13 or 14, wherein agitating the substantially aqueous solution comprises at least one of stirring the solution, recirculating the solution, injecting the solution through one or more openings in the catalyst, and treating the solution with ultrasonic or low frequency vibration.

16. The method of any of claims 1 to 15, further comprising subjecting the catalyst to a pretreatment, wherein the pretreatment comprises one or more of a mechanical pretreatment to remove at least a portion of fly ash from a surface of the catalyst, washing the catalyst with water or a second aqueous solution, and combinations thereof.

17. The method of any of claims 1 to 16, further comprising washing the SCR catalyst with water or a third aqueous solution after the treating step and drying the SCR catalyst.

18. The method of any of claims 1 to 17, further comprising re-impregnating the SCR catalyst with one or more catalytic active metal compounds.
19. The method of any of claims 1 to 18, wherein the treated SCR catalyst has a decreased
$\text{SO}_2$ to $\text{SO}_3$ conversion rate compared to the untreated SCR catalyst.

20. A method of regenerating a DeNOx catalyst having an accumulation of one or more iron compounds, the method comprising:
   - treating a DeNOx catalyst having an accumulation of one or more iron compounds
   - with a substantially aqueous solution having a pH ranging from about 2.0 to about 5.0 and
   - comprising from about 0.1% to about 4.0% by weight of an acidic reactive ammonium salt
   and from about 0.5% to about 2.0% by weight of ascorbic acid or derivatives thereof;
   - removing at least a portion of the one or more iron compounds accumulated on the
   - DeNOx catalyst;
   - washing the treated DeNOx catalyst with water or an aqueous solution; and
   - drying the DeNOx catalyst.

21. The method of claim 20, wherein the treated DeNOx catalyst has a decreased $\text{SO}_2$ to $\text{SO}_3$ conversion rate compared to the untreated DeNOx catalyst.
FIG. 1a

PH-Value of Ammonium Salts at Different Temperatures and Varying Ascorbic Acid Concentration

20 deg C  60 deg C  60 deg C + 0.5%wt Ascorbic Acid  60 deg C + 1.0%wt Ascorbic Acid  60 deg C + 2.0%wt Ascorbic Acid

Test Condition
FIG. 2

Total Iron-Trendline during Regeneration with Hydrofluoric Acid and Ammonium Salts in Combination with Ascorbic Acid

Test Condition

Room Temperature / 0hr  50 deg C / 1hr  50 deg C / 4hr

Total Iron Concentration [mg/l]

- HF 0.05% / AA 0.5%
- HF 0.1% / AA 0.5%
- HF 0.15% / AA 0.5%
- HF 0.2% / AA 0.5%
- NH4Cl 0.5% / AA 0.3%
- NH4Cl 0.5% / AA 0.5%
- NH4Cl 1% / AA 0.5%
- NH4Cl 2% / AA 0.5%
- (NH4)2SO4 0.5% / AA 0.3%
- (NH4)2SO4 0.5% / AA 0.5%
- (NH4)2SO4 1% / AA 0.5%
- (NH4)2SO4 2% / AA 0.5%