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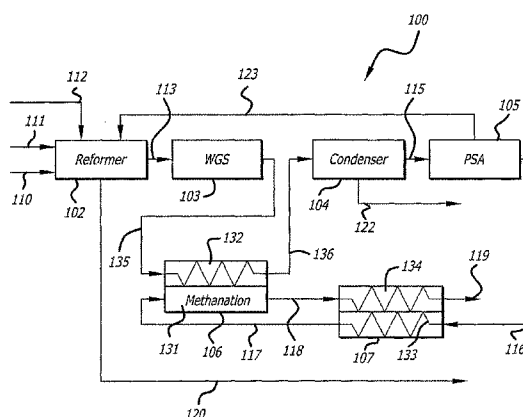
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(54) Title: HYDROGEN PURIFICATION PROCESS AND SYSTEM



(57) Abstract: A hydrogen generation system is disclosed that includes a fuel reforming reactor generating a hydrogen-rich reformat gas at a temperature greater than 150C, a pressure swing adsorption (PSA) hydrogen purification unit that separates the reformat gas into a relatively pure hydrogen stream and an off-gas stream, and a catalytic reactor down stream of the PSA unit that converts carbon monoxide (CO) and hydrogen (H₂) contained in the relatively pure hydrogen stream into methane (CH₄) and water vapor (H₂O). The method of purification involves generating a hydrogen-rich reformat gas at a temperature greater than 150C in a fuel reforming reactor, separating the reformat gas into a relatively pure hydrogen stream and an off-gas stream in a pressure swing adsorption (PSA) hydrogen purification unit, and converting carbon monoxide (CO) and hydrogen (H₂) contained in the relatively pure hydrogen stream into methane (CH₄) and water vapor (H₂O) in a catalytic reactor down stream of the PSA unit. The hydrogen can be further purified by including a secondary purification stage downstream of the PSA unit and the catalytic reactor wherein the secondary purification stage has a water adsorbent material bed that adsorbs the water vapor H₂O and a hydrogen absorbent material downstream of the water adsorbent material that absorbs hydrogen gas preferentially, thus concentrating the non-hydrogen components, such as CH₄, into an exhaust stream that exits the bed, wherein the absorbed hydrogen gas is then desorbed to create an exiting very pure hydrogen stream.



WO 2007/106139 A2



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HYDROGEN PURIFICATION PROCESS AND SYSTEM

BACKGROUND

1. Field

This disclosure relates to improving the purity of hydrogen gas using a secondary hydrogen purification method downstream of a primary hydrogen pressure swing adsorption unit.

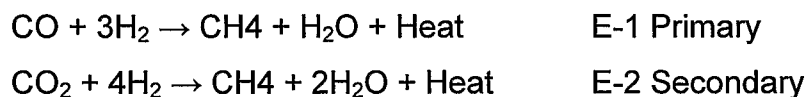
2. General Background

A hydrogen generation unit (HGU) is a combination of thermo-chemical processes that convert a fuel-steam mixture into a hydrogen-rich gas mixture typically composed of hydrogen (H_2), carbon monoxide (CO), carbon dioxide (CO_2), methane (CH_4), water vapor (H_2O) and other gases depending on the composition of the fuel feedstock. Typically this mixture is known as reformat. For many applications this reformat stream is then passed to a hydrogen purification unit in which 60% to 90% of the hydrogen is separated into a relatively pure hydrogen stream (99+% H_2) and an off-gas stream consisting of the other species in the reformat mixture. One typical method used to purify the reformat is a pressure swing adsorption (PSA) unit, which consists of a series of bed filled with adsorbent material (typically but not limited to zeolites). As the pressurized reformat flows through the bed gaseous species adsorb on to the active surfaces. Since the H_2 is the least strongly adsorbed species in the reformat stream, a pure H_2 gas exits the bed. After a period of time when the adsorbent sites begin to become saturated, the feed gas is removed and the bed is depressurized forcing the adsorbed species to desorb and exit the bed as the off-gas stream. By cycling several beds through this pressurization and depressurization cycle a continuous H_2 purification process is created. As the capacities of the beds are pushed to their limits with higher flow rates and faster cycle times, non-hydrogen gas species

begin to contaminate the relatively pure H₂ gas stream. Typically, the species of concern are the other gases in the reformat stream such as CH₄, CO, and CO₂.

Of the non-hydrogen species typically in the reformat feed to the PSA, H₂O and CO₂ are strongly adsorbed onto the surfaces of the zeolites and CO and CH₄ are weakly adsorbed. As a result the relatively pure hydrogen stream exiting the PSA typically has CO and CH₄ as the primary contaminants. In fuel cell and hydrogen refueling station applications the most critical of these contaminants is the CO, because it causes performance degradation of the fuel cell or metal hydride hydrogen storage units. CH₄ is relatively non-reactive in the fuel cell and metal hydride materials, and therefore, does not cause performance degradation. It is beneficial to include a reactor between the PSA and the fuel cell that converts the CO back into CH₄. This reactor allows the capacity of the PSA to be increased substantially without impacting performance of fuel cell units downstream.

One reaction mechanism that achieves this is known as methanation, which is the reverse of the steam reforming reactions. Typically in methanation reactors, a catalyst is used and ruthenium based catalysts have proven to be very effective, although other catalysts such as nickel, platinum, etc. can be used. For these catalysts to be effective the temperature of the catalyst must be greater than 150°C and preferably greater than 190°C.



Methanation reactors have been integrated into hydrogen generation systems in the prior art, but typically they have been used upstream of the PSA unit to minimize the CO concentration entering the PSA unit. Typically, this upstream location is used because the reformat gas is at the appropriate temperature range to maintain the catalyst activity. One issue with this art is the secondary reaction identified above, where CO₂ is also converted into CH₄ since

the CO₂ concentration is typically 20% in comparison to CO concentrations in the 2 to 4% range. To manage the methanation reaction a tight temperature range is preferably maintained in the catalytic bed which promotes the reaction of CO but does not promote the reaction of the secondary reaction with CO₂. Not only is this a source of process inefficiency, but it can also result in thermal run-away in which all the CO₂ is reconverted back into CH₄. Therefore a system and method are needed which alleviates these inefficiencies and precludes thermal run-away from occurring.

SUMMARY OF THE DISCLOSURE

The present disclosure is directed to a hydrogen generation system and method that fundamentally includes operations of generating a hydrogen-rich reformat gas at a temperature greater than 150C in a fuel reforming reactor, separating the reformat gas into a relatively pure hydrogen stream and an off-gas stream in a pressure swing adsorption (PSA) hydrogen purification unit, and converting low level concentrations of carbon monoxide (CO) contained in the relatively pure hydrogen stream into methane (CH₄) and water vapor (H₂O) in a catalytic reactor down stream of the PSA unit. Preferably the hydrogen generation method also includes positioning the catalytic reactor such that the temperature of the hydrogen-rich reformat gas maintains the temperature of the catalytic reactor.

Alternatively, or, in addition, the hydrogen generation system and method includes integrating the fuel reforming reactor with a combustion reactor that uses the off-gas stream to provide thermal energy to the fuel reforming reactor. Here the combustion reactor has a combustion exhaust gas at a temperature greater than 150C, and includes positioning the catalytic reactor such that the temperature of the exhaust gas maintains the temperature of the catalytic reactor.

The hydrogen generation method and system may include a secondary purification stage downstream of the PSA unit and the catalytic reactor. The secondary purification stage has a water adsorbent material bed that adsorbs the H₂O vapor; and concentrates non-hydrogen components into an exhaust stream exiting the bed by providing a hydrogen absorbent (or adsorbent) material downstream of the water absorbent material bed. Finally, the absorbed hydrogen gas is desorbed to generate a very pure hydrogen stream.

The above-mentioned features and objects of the present disclosure will become more apparent with reference to the following description taken in conjunction with the accompanying drawings.

DRAWINGS

Figure 1 illustrates one embodiment of a system of the disclosure in which thermal energy of the reformat stream is used to maintain the catalyst at optimum conditions.

Figure 2 illustrates a second embodiment of the system in which thermal energy of the combustion exhaust gas is used to maintain the catalyst at optimum conditions.

Figure 3 is a graph that illustrates test data collected on a hydrogen generation unit shown in Figure 1.

Figure 4 illustrates an exemplary system in which a secondary hydrogen purification process is added downstream of the methanation reactor in either of the systems shown in Figures 1 or 2.

Figure 5 is a set of tables showing predicted hydrogen production levels at various stages in the systems shown in Figures 1, 2, and 4.

Figure 6 is a graph of critical CO conversion and hydrogen purity in the purification process shown in Figure 4.

DETAILED DESCRIPTION

Figure 1 illustrated one exemplary embodiment of a system in accordance with the disclosure in which a methanation reactor is integrated into a system based on steam reformation and PSA purification. Basically the methanation reactor is integrated into the system such that the reformat gas stream exiting the water gas shift (WGS) reactor at a temperature greater than 150C is used as the thermal energy source to maintain the methanation reactor at its optimum temperature. Typically the exit temperature of the WGS exit stream is in the range of 150C to 400C, and more preferably in the range of 200C to 300C.

The embodiment shown in Figure 2 is similar to that of Figure 1 except that combustion exhaust gas is used to maintain temperature in the methanation reactor. Both of these embodiments provide the unique advantage of the disclosure that places the methanation reactor downstream of the PSA. Since most if not all the CO₂ is removed from the relatively pure hydrogen product stream from the PSA, there is no need for tight temperature control because the concentration of reactants (CO and CO₂) are relatively low (typically below 1000 ppm) and the purification goal is to convert all these reactant species into CH₄. Therefore, there is no upper limit to the catalytic reactor's temperature, because the concentration of CO₂ in the relatively pure hydrogen stream are insufficient to allow thermal runaway of the methanation reactor.

Now specific reference is made to the first embodiment of the system shown in Figure 1. This embodiment of the hydrogen generation system 100 consists of a reformer 102, WGS reactor 103, condenser 104, PSA 105, methanation reactor 106, and hydrogen gas recuperative heat exchanger 107. The reformer 102 has an integrated combustor section and steam reforming section

(not detailed). Each reactor section or process in the reformer **102** has an optimum temperature range based on the specific catalysts and processes that are occurring.

5 The combustion section receiving air **112** and PSA off gas **123** which is typically 40 to 60% hydrogen reacts this fuel air mixture generating heat and operating at a temperature in the 700 to 950C range. The reformer **102** typically operates in the 650 to 900C range or at a temperature 50 to 100C cooler than the combustion section within the reformer **102** to promote the transfer of heat from the combustion section to the reformation section. The WGS typically operates in
10 the 500 to 250C range depending on the specific catalysts. The PSA's temperature range is typically in the 40 to 80C range, and requires no liquid water in the feed. Therefore the condenser **104** is used to cool and dry the reformat stream **136** and can include a reheat function to ensure that the reformat stream **115** entering the PSA **105** has a dew point less than the dry bulb temperature. The
15 ideal methanation reactor temperature is above 190C. Since the PSA operating temperature is typically lower than the methanation temperature, methanation reactors are typically not utilized downstream of the PSA.

Fuel **110** and water vapor or steam **111** enters the steam reforming section of the reformer **102** where the steam-fuel mixture is converted into a pre-WGS
20 reformat stream **113** containing H₂, CO, CO₂, H₂O, CH₄, and other trace species. This stream **113** enters the WGS **103** where additional CO is reacted with H₂O to form H₂ and CO₂ creating the reformat stream **135** that exits the WGS **103**. In some alternative embodiments the reformer **102** and WGS **103** can be integrated into a single assembly, and in others the WGS can be minimized or even
25 eliminated. The thermal energy contained in the reformat stream **135** is used to heat the methanation reactor **106** by indirect transfer from heat exchanger surface **132** to methanation catalyst **131**. The reformat steam **136** flows to the condenser

104 in which thermal energy is removed to manage the inlet PSA **105** temperature and to condenser water from the reformat mixture. The condensate water **122** is removed. The drier reformat gas **115** is passed to the PSA **105** in which it is separated into a relatively pure hydrogen stream **116** and an off-gas stream **123**.

5 The relatively pure hydrogen stream **116** is raised in temperature by the recuperative heat exchanger **107**. The temperature of the relatively pure hydrogen stream **116** is dependant on the sizing and configuration of the recuperative heat exchanger **107**, but typically temperatures in the 100C to 150C are achievable depending on the temperature of the methanation reactor. The hotter relatively
10 pure hydrogen stream **117** then enters the methanation reactor **106** in which any CO and CO₂ contaminates are reacted with H₂ to form CH₄ and H₂O. The temperature of the methanation reactor **106** is managed by the indirect heat transfer with the reformat stream **135** and the heat exchange surface **132**. Although the methanation reactions (E-1 and E-2) are exothermic, the heat
15 generated is not sufficient to control the temperature of the methanation reactor **106** because the concentrations of CO and CO₂ in the feed stream **117** are relatively low, typically less than 1000 ppm or 0.1%. The low CO concentration, relatively pure hydrogen stream **118** exits the methanation reactor **106** and provides thermal energy to the recuperative heat exchanger **107**. Typically the
20 temperature of the exit hydrogen stream **119** is in the range of 100 to 160C depending on the surface area and design of heat exchanger **107**.

 The advantage of this system **100** and the method of its operation is to produce a relatively pure hydrogen exit stream **119** in which the only impurities are CH₄, H₂O and other non-reactive species, such as nitrogen, argon, etc. The
25 reactive species CO and CO₂ have been converted back into CH₄, which is non-reactive at the operating temperatures of a proton exchange membrane (PEM) fuel cell. This relatively pure hydrogen exit stream **119**, for example, can

be sent directly to a PEM fuel cell for power generation. The PEM fuel cell consumes the hydrogen to generate power and the CH₄ and other non-reactive species will build up concentration in the dead headed flow path. Periodically a valve will open forcing the collected CH₄ and other non-reactive species out of the fuel cell to maintain performance while achieve 99% hydrogen utilization in the fuel cell.

Figure 2 illustrates a second embodiment **200** of the system in accordance with the present disclosure. In this embodiment **200** like numbers are used as in Figure 1 for the same components. Embodiment **200** differs specifically from system **100** in that exhaust gas **120** from the reformer **102** feeds directly into and through the heat exchanger surface **132** rather than the discharge of the WGS **103** feeding the heat exchanger **132**. Thus the thermal energy used to maintain the temperature of the methanation reactor **106** is energy from the exhaust gas **120** from the combustion section of the reformer **102**. Since this temperature source is used, the reformat stream **214** from the WGS **103** flows directly into the condenser **104**. The combustion exhaust **120** from the reformer **102** is typically used for feed air **112** preheating and exits the reformer **102** in the temperature range of 150 to 300C depending on the preheating characteristics.

Figure 3 illustrated actual test data graph **300** collected from a HGU with a methanation reactor in accordance with the present disclosure. The HGU was an integrated steam reformer and water gas shift reactor system **100** as illustrated in Figure 1. The reformat gas **135** was used to provide the thermal energy for the methanation reactor **106** prior to flowing to the condenser **104** and the PSA unit **105**. The CO and CH₄ concentration after the PSA unit **105** and after the methanation reactor **106** are plotted. The relatively pure hydrogen stream **119** exiting the test hardware had a CO concentration of less than 0.5ppm as illustrated by the open square data set **341** and a CH₄ concentration as illustrated

by the downwardly pointed open triangle data set **344**. Both data sets are plotted with hydrogen purity as the x-axis. The open circle data set **342** illustrates the CO concentration in stream **116** exiting the PSA **105** and the upwardly pointed open triangle data set **343** illustrates the CH₄ concentration in stream **116**. This data set indicates the increase in CH₄ concentration after the methanation reactor **106** which corresponds to the decrease in CO concentration.

In some refueling station and industrial gas applications a pure hydrogen product of greater than 99.9% is required. To achieve this level of purity, typically the PSA unit is increased in size and the hydrogen recover in the PSA unit is decreased until there are very few non-hydrogen species in the PSA product gas **116**. A hydrogen concentration of 99.99+% is equivalent to less than 100ppm non-hydrogen concentration and a 99.999+% is equivalent to less than 10ppm non-hydrogen concentration. This requirement can force the PSA unit **105** to be large and costly.

The systems **100** and **200** illustrated in Figure **1** and Figure **2** can also be used in these applications with an additional secondary purification system **400** positioned downstream of the relatively pure hydrogen exit stream **119**. The secondary purification system **400** allows the primary PSA unit **105** to be decreased in size and cost and operated at a faster cycle time, which also helps decrease pressure pulsations in both the off-gas and product gas. The secondary purification subsystem **400** is illustrated in Figure **4**.

The objective of subsystem **400** is to purify the reformate stream with a PSA unit **105** allowing CH₄, CO, and CO₂ as contaminants at concentrations less than 0.1% or 1000 ppm. This allows the PSA unit **105** to be considerably downsized and a very rapid cycle can be used to maximize the capacity of the PSA unit **105**. Downstream of the PSA unit **105** a methanation reactor **106** as indicated in Figures **1** or **2** is used to convert the CO and CO₂ into additional CH₄

and H₂O. Downstream of the methanation reactor **106** this hydrogen stream is cooled and passed through a desiccant material to remove the water vapor produced in the methanation reactor **106**. This dry stream is then passed into a H₂ absorbent bed such as a metal hydride in which the H₂ is absorbed into the material and stored for later release.

Removal of the water vapor and CO is important in this secondary purification subsystem **400**, because these species can damage the performance and endurance of the metal hydride bed. The absorbent bed is designed to have a high length over width ratio such that as the hydrogen gas is adsorbed the CH₄ and other non-reactive, non-condensable species are concentrated as they move down the bed. Finally, the gases exit the bed as a hydrogen-rich mixture with higher concentrations of CH₄ and other non-reactive species. These gases are vented from the secondary purification system **400** and may be returned to the combustion section of the reformer **102**. Valves in the system **400** are used to isolate the beds and to open the H₂ absorbent to the product H₂ supply line. The absorbed H₂ is desorbed resulting in a very pure H₂ stream, typically in the 99.995+% range because the only non-hydrogen is the contaminants CH₄ and other non-reactive, non-condensable species in the gas volume around the absorbent material.

In reference to Figure 4 one embodiment of the secondary hydrogen purification system **400** consists of four beds, first desiccant bed **401**, first H₂ absorbent bed **402**, second desiccant bed **403** and second H₂ absorbent bed **404**. The system **400** also has a series of valves that manage the flow of relatively pure hydrogen stream from the methanation reactor **106** and feed line **440** to the product hydrogen supply line **442** and the vent line **441**. In this embodiment first and second parallel paths **408** and **409** are illustrated which cycle on and off in sequence to create a continuous process. Other embodiments may include three

or more parallel paths to optimize purity and process integration. The open valves of the first parallel path **408** are shown as white or open symbols, while the closed valves of the second parallel path **409** are shown as dark or solid symbols. The valves are shown as on-off solenoid valves, but could also be rotary valve assemblies similar to the valves used in a rapid cycle PSA unit. Rotary valves assemblies provide advantages of reliability, lower cost, and simple process control.

Process gases from the methanation reactor **106** flow into the system **400** through supply line **440**. In Figure **4** the first parallel path **408** supply valve **431** is open allowing the hot gas from the methanation reactor **106** to enter the first heat exchanger coil **410** of first desiccant bed **401**, while second parallel path supply valve **451** is closed. The hot gas from the methanation reactor **106** transfers heat to the desiccant material **412** of desiccant bed **401**. The gas flows from coil **410** to second heat transfer coil **411** embedded in first H₂ absorbent bed **402** again transferring heat to the first H₂ absorbent material **413**. Absorbed H₂ from the prior cycle in the first material **413** is desorbed and flows through open first exit valve **436** and out of the system through product line **442**.

H₂ desorption is endothermic which decreases the temperature of the first absorbent material **413** and helps transfer heat from second coil **411** cooling the gas temperature in coil **411**. The gas flows from the second coil **411** through open valve **132** and into second desiccant bed **403** and past the second desiccant material **462** which adsorbs any water vapor in the gas stream generated by the methanation reactions in bed **106**. The dry hydrogen rich gas flows through open valve **433** and into second H₂ absorbent bed **404** and H₂ absorbent material **463**. The hydrogen in the stream is absorbed as the gas flow through second absorbent bed **463** thereby increasing the concentration of CH₄ and other non-reactive species in the process flow. This gas then exits the second H₂ absorbent bed **404**

through open valve **434** and into first desiccant bed **401** in contact with first desiccant material **412**. The heat transferred from first coil **410** causes the adsorbed water vapor in first desiccant material **412** to evolve and flow out of the system through open valve **435** to system vent **441**.

5 Once the second desiccant material **462** or the second absorbent material **463** nears saturation the process cycle is reversed and flow is diverted through the second parallel path **409** by closing the open valves **431**, **432**, **433**, **434**, **435**, and **436** and opening the closed valves **451**, **452**, **453**, **454**, **455**, and **456**. When the secondary parallel path **409** is active, heat transferred in third coil **457** and fourth
10 coil **458**, while water vapor is adsorbed in first desiccant material **412** and hydrogen is absorbed in first absorbent material **413**. Similarly, product hydrogen is desorbed from second absorbent material **463** flowing through valve **456** toward product line **442**.

15 In one embodiment of the secondary purification system **400** the hydrogen absorbent material **413** and **463** is a metal hydride type material. The material absorbs H₂ molecules into the metal lattices and has the capacity to absorb approximately 150 to 300 times the volume of hydrogen at standard conditions in comparison to the volume of metal. Any absorbent or adsorbent materials can be used which preferentially remove H₂ gas from a mixture.

20 The method described is very simple in that all the open valves are closed and all the closed valves are open uniformly. More complex and potentially more effective cycles are embodied. For example the purge valve **434** positioned between the second H₂ absorbent material **463** and the first desiccant material **412** can be opened after the starting the process parallel path **408**, to maintain the
25 pressure of the process gases in absorbent material **463** and thereby, enhance the percentage of the process hydrogen gas in the flow stream that is absorbed in the material **463**.

Similarly, the open valves of parallel path **408** can be closed except for purge valve **434** and exhaust valve **435** which are left open to allow a brief period during which process gases in the gas volumes of absorbent material **463** are purged due to depressurization and desorption of pure hydrogen. The pure hydrogen pushes any process gas with contaminants out through valve **434**. After the brief period parallel path **463** valves are closed and process flow is initiated through parallel path **409**. As a result the hydrogen gas allowed to flow through product valves **436** and **456** is extremely pure achieving purities one magnitude greater than without this delay.

An analytical model was developed to assess the performance characteristics and sensitivities of the process to the various parameters. One set of results from this model is shown in Figure 5. This Figure 5 is divided into six blocks of data. The first block (Table 5.A) shows a typical reformat composition entering the PSA primary hydrogen purification unit. The analysis was conducted for one multiple of 100 moles at this point, and the concentrations of gas species are shown in ppm and percent in the third and fourth column of each block of data, respectively.

The second block (Table 5.B) illustrates an assumed relatively pure hydrogen stream **116** leaving the primary purification unit **105** as in Figure 1. Assuming that 70% of the hydrogen was recovered, the CO concentration was set at 200ppm and the CH₄ concentration was set at 800ppm, with 10ppm of CO₂ and 3ppm or other non-reactive species such as argon or nitrogen. This composition is representative of the relatively pure hydrogen stream **116**. Assuming the initial 100 moles of gas and 74 moles of H₂ the relatively pure hydrogen stream has 51.8 moles of H₂ at the 70% recovery rate. The hydrogen purity of this stream is only 99.899%, which is acceptable for a fuel cell from a bulk characteristic except for the 200ppm of CO which would degrade PEM fuel cell performance very rapidly.

To address this issue a methanation reactor **106** is integrated into the system (Figures **1** or **2**). The methanation reaction converts 99.975% of the CO and CO₂ into CH₄ and H₂O and the outlet conditions are shown in the third block of data (Table 5.C). The H₂ concentration has decreased slightly because of the H₂ losses, but the output of the methanation reactor **106** is acceptable for direct fuel cell consumption because the CO concentration is 0.05ppm well below the required 1 ppm level. Actually if only 99.5% of the CO is converted in the methanation reaction the CO concentration would be 1 ppm in the outlet.

The water content after the methanation reaction is 220ppm, and because the assumed H₂ absorbent is a metal hydride, this concentration of water would damage the hydride. The desiccant material removes this water in a process assumed to be 99.99% effective as shown (Table 5.D). As the gas stream flow through the hydride, H₂ is absorbed. If 95% of the hydrogen is absorbed, the outlet concentration of CH₄ and CO has increased as shown (Table 5.E). The CO level is still approximately 1 ppm and the CH₄ concentration has increased from 1010 ppm to 19830 ppm or just under 2%. Since only 95% of the hydrogen is absorbed, the hydrogen concentration is still 98% and 2.59 moles of H₂ leave the secondary system with the CH₄ and CO. The 49.18 moles of H₂ absorbed results in the net H₂ recovery being 66.5% (94.9% of the initial 70% value). Assuming that the metal hydride had only a 150 times H₂ gas volume to vessel volume, when the desorbed H₂ during the next parallel path cycle this volume of H₂ is mixed with one volume of absorbent bed inlet gas (with 1010ppm CH₄ and H₂ concentration of 99.9%). The result is an H₂ purity of 99.9993% after the secondary purification process as shown (Table 5.F).

One aspect of the durability of the secondary purification process **400** is the concentration of CO at the end of the H₂ absorbent beds **413** or **463** as the H₂ is absorbed. An example of this analysis is presented in Figure **6** illustrating the CO

concentration at the end of the absorption bed as a function of the percent of H₂ recovered in the secondary bed (x-axis) and the level of CO conversion in the methanation reactor **106**. With decreased activity in the methanation reactor (**106**) the H₂ recovery in the secondary purification process **100** must be decreased to
5 maintain CO concentrations at an acceptable level. If H₂ absorbent materials **413** and **463** with higher tolerance to CO are used optimization of these parameters can be adjusted. Higher activity of the methanation reactor **106** is achieved by increasing the temperature of the catalyst. The innovations defined in Figures **1** and **2** are needed to passively achieve reactor **106** temperatures without too much
10 system complexity or process inefficiencies.

While the apparatus and method have been described in terms of what are presently considered to be the most practical and preferred embodiments, it is to be understood that the disclosure need not be limited to the disclosed
embodiments. It is intended to cover various modifications and similar
15 arrangements included within the spirit and scope of the claims, the scope of which should be accorded the broadest interpretation so as to encompass all such modifications and similar structures. The present disclosure includes any and all embodiments of the following claims.

CLAIMS

1. A hydrogen generation method comprising:
generating a hydrogen-rich reformat gas at a temperature greater
than 150C in a fuel reforming reactor;
5 separating the reformat gas into a relatively pure hydrogen stream
and an off-gas stream in a pressure swing adsorption (PSA) hydrogen purification
unit, and
converting carbon monoxide (CO) and hydrogen (H₂) contained in
the relatively pure hydrogen stream into methane (CH₄) and water vapor (H₂O) in
10 a catalytic reactor down stream of the PSA unit.
2. The hydrogen generation method of claim 1 further comprising:
positioning the catalytic reactor such that the temperature of the
hydrogen-rich reformat gas maintains the temperature of the catalytic reactor.
3. The hydrogen generation method of claim 1 further comprising:
15 integrating the fuel reforming reactor with a combustion reactor that
uses the off-gas stream to provide thermal energy to the fuel reforming reactor, the
combustion reactor having a combustion exhaust gas at a temperature greater
than 150C; and
positioning the catalytic reactor such that the temperature of the
20 exhaust gas maintains the temperature of the catalytic reactor.
4. The hydrogen generation method of claim 1 further comprising:
providing a secondary purification stage downstream of the PSA unit
and the catalytic reactor, the secondary purification stage having a water
adsorbent material bed that adsorbs the H₂O; and

concentrating non-hydrogen components into an exhaust stream exiting the bed by providing a hydrogen absorbent (or adsorbent) material downstream of the water absorbent material bed; and

5 desorbing the absorbed hydrogen gas to generate a very pure hydrogen stream.

5. A hydrogen generation system comprising:

a fuel reforming reactor generating a hydrogen-rich reformat gas at a temperature greater than 150C;

10 a pressure swing adsorption (PSA) hydrogen purification unit that separates the reformat gas into a relatively pure hydrogen stream and an off-gas stream, and

a catalytic reactor down stream of the PSA unit that converts carbon monoxide (CO) and hydrogen (H₂) contained in the relatively pure hydrogen stream into methane (CH₄) and water vapor (H₂O).

15 6. The hydrogen generation system of claim 5 further comprising:

the catalytic reactor being positioned such that the temperature of the hydrogen-rich reformat gas maintains the temperature of the catalytic reactor.

7. The hydrogen generation system of claim 5 further comprising:

20 a combustion reactor integrated with the fuel reforming reactor that uses the off-gas stream to provide thermal energy to the fuel reforming reactor;

the combustion reactor having a combustion exhaust gas at a temperature greater than 150C exiting the combustion reactor; and

wherein the catalytic reactor is positioned such that the temperature of the exhaust gas maintains the temperature of the catalytic reactor.

25 8. The hydrogen generation system of claim 5 further comprising:

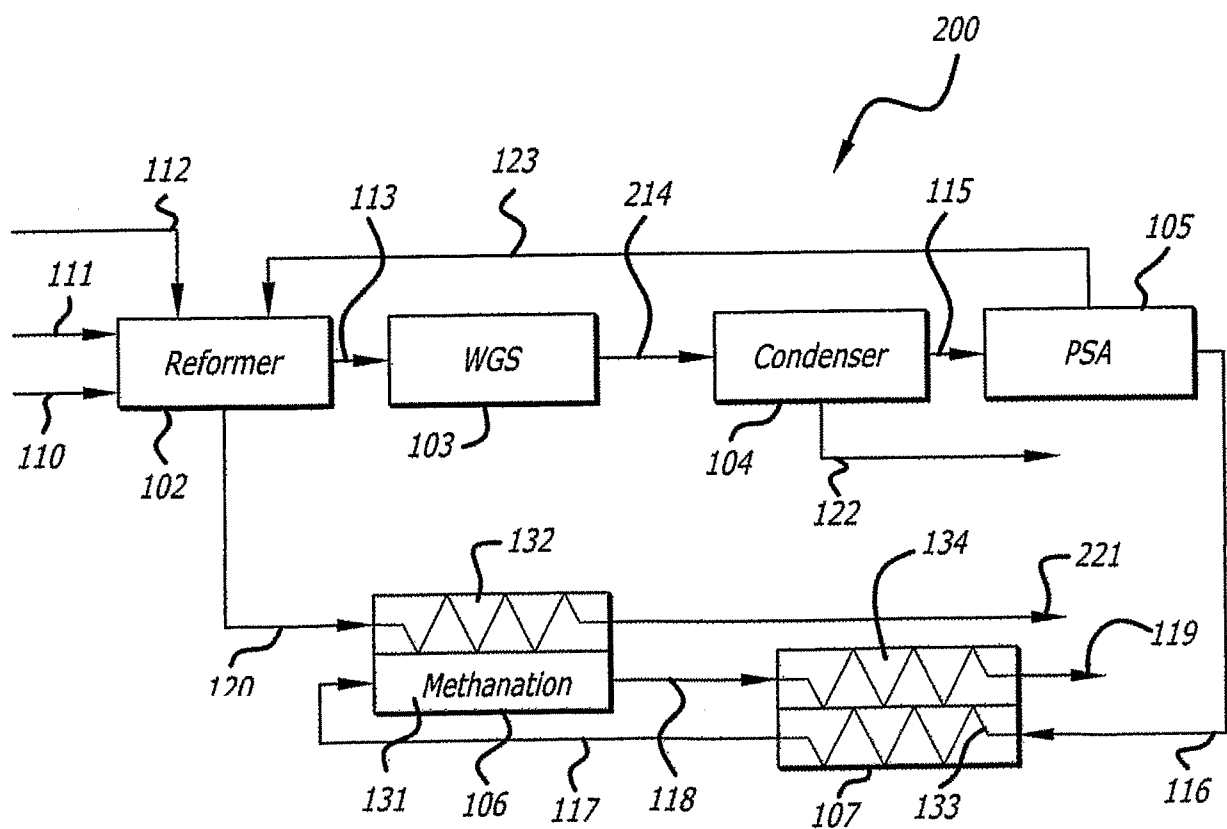
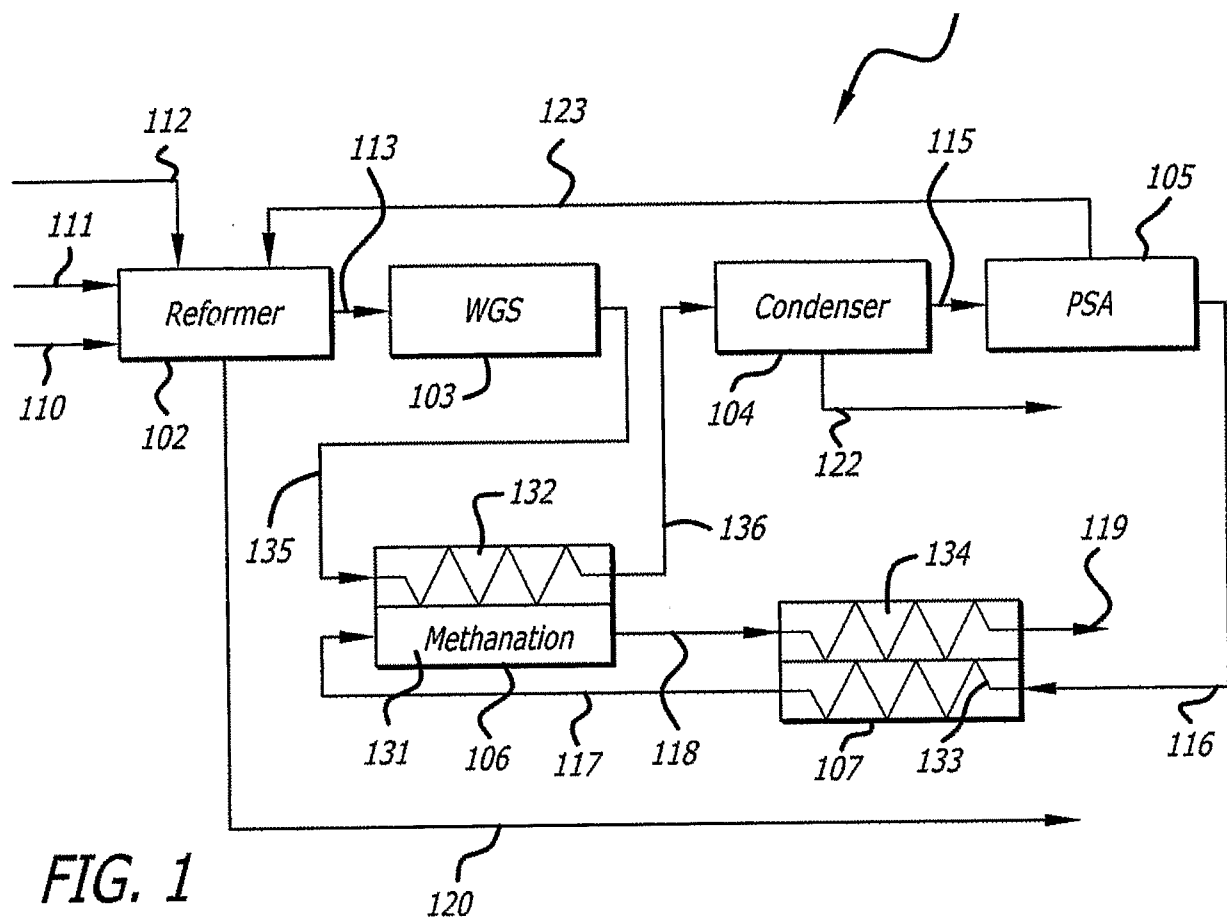
a secondary purification stage downstream of the PSA unit and the catalytic reactor wherein the secondary purification stage has a water adsorbent material bed that adsorbs the water vapor H_2O ; and

5 a hydrogen absorbent material downstream of the water absorbent material that absorbs hydrogen gas preferentially, thus concentrating the non-hydrogen components, such as CH_4 , into an exhaust stream that exits the bed, wherein the adsorbed hydrogen gas is then desorbed to create an exiting very pure hydrogen stream.

9. The hydrogen generation system of claim 5 further comprising:

10 a secondary purification stage downstream of the PSA unit and the catalytic reactor wherein the secondary purification stage has a water adsorbent material bed that adsorbs the water vapor H_2O ; and

15 a hydrogen absorbent material downstream of the water absorbent material that absorbs hydrogen gas preferentially, thus concentrating the non-hydrogen components, such as CH_4 , into an exhaust stream that exits the bed, wherein the adsorbed hydrogen gas is then desorbed to create an exiting very pure hydrogen stream.



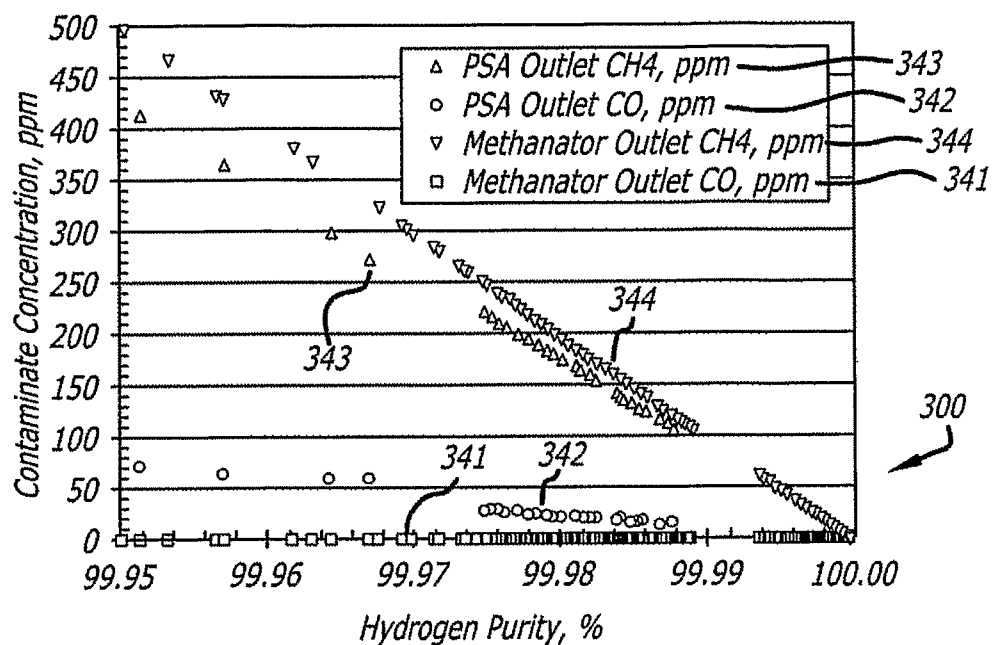


FIG. 3

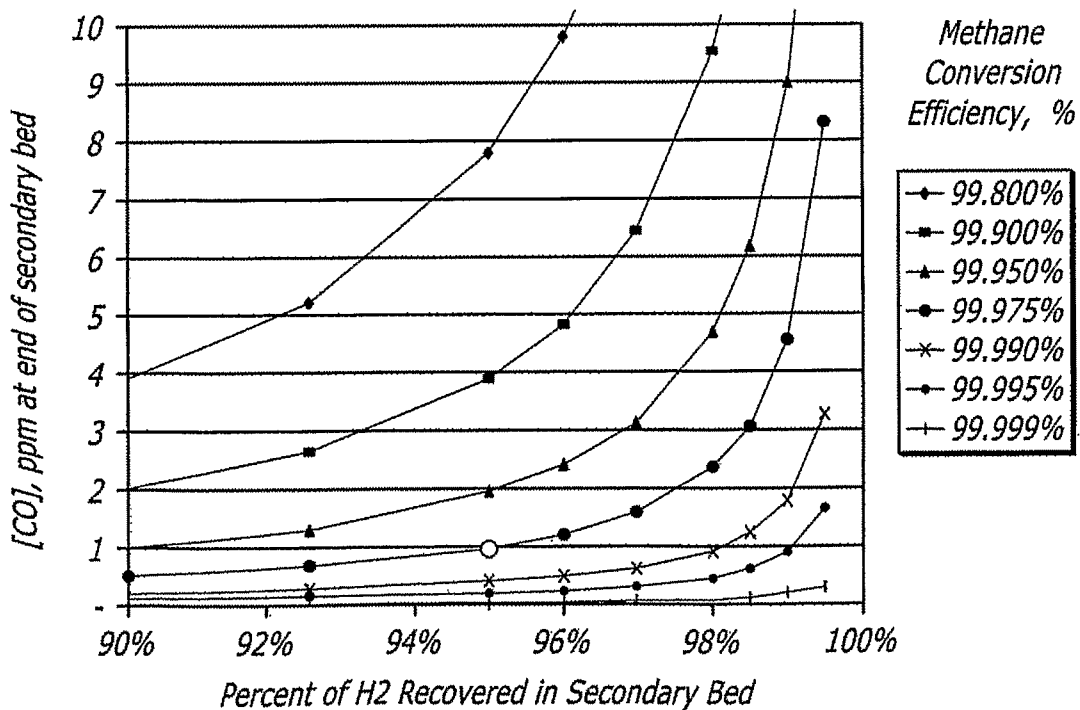


FIG. 6

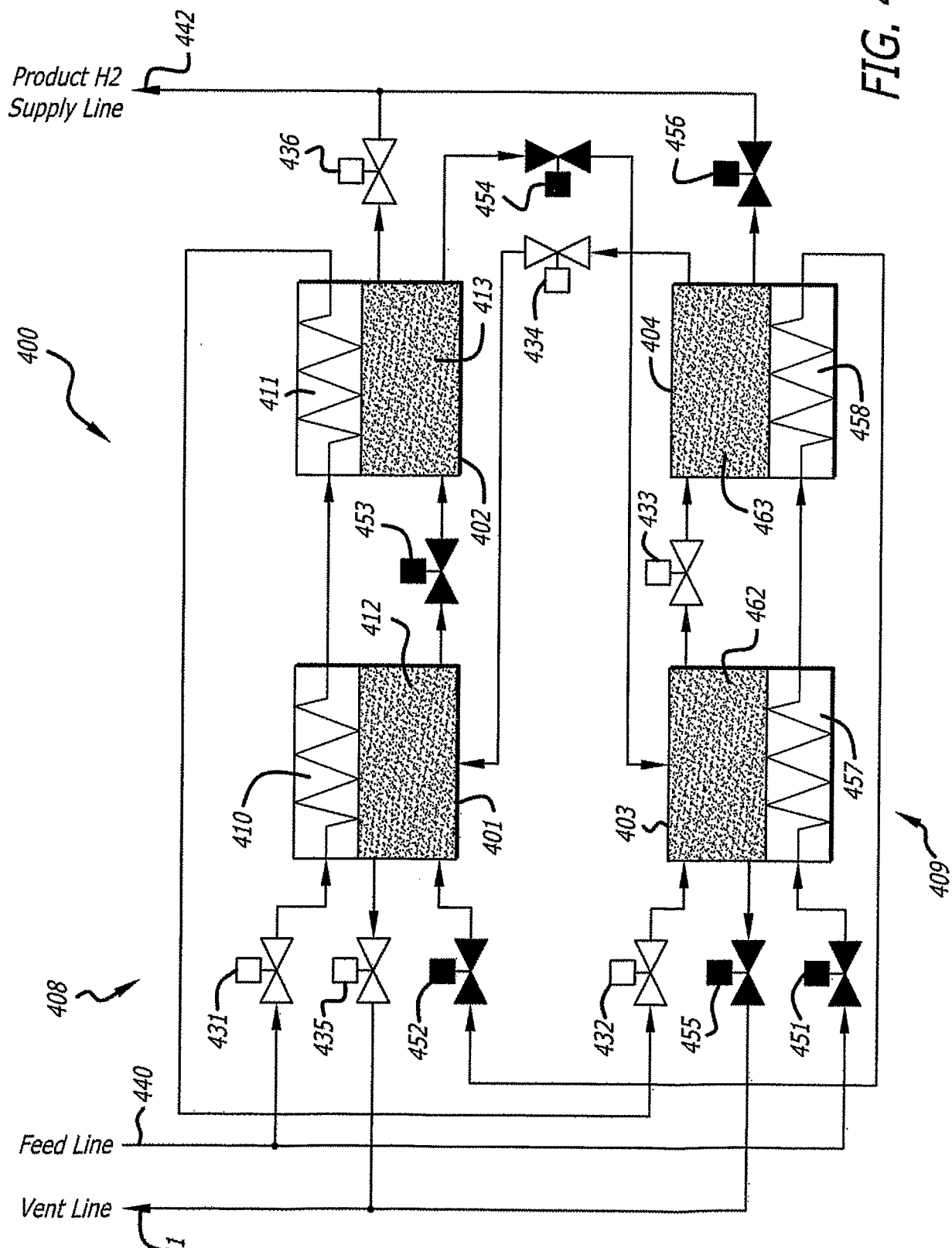


FIG. 4

Table 5.A Inlet Reformate Composition

Assumption: Typical Liq. HC Reformate

	moles	ppm	%
H2	74	740000	74.0
CO	2	20000	2.0
CO2	21	210000	21.0
CH4	2.9	29000	2.9
H2O			
Other	0.1	1000	0.1
moles	100		
%		100.0	100.0

Table 5.D After H2O Absorption

Assumption: 99.990% Adsorption

	moles	ppm	%
H2	51.7668	998986.3	99.899
CO	0.0000	0.051	0.000
CO2	0.000000	0.003	0.000
CH4	0.0524	1010.6	0.101
H2O	0.0000	0.022	0.000
Other	0.0002	3.002	0.000
moles	51.82		
%		100.0	

Table 5.B After PSA

Assumption: 70.0% PSA H2 Recovery

	moles	ppm	%
H2	51.8000	998987	99.899
CO	0.0104	200	0.020
CO2	0.0005	10	0.001
CH4	0.0415	800	0.080
H2O	0.0000	0	0.000
Other	0.0002	3	0.000
moles	51.85		
%	51.8525	100.0	

Table 5.E After H2 Absorption

Assumption: 95.0% H2 Absorption

	moles	ppm	%
H2	2.59	980109.7	98.011
CO	0.0000	1.000	0.000
CO2	0.0000	0.050	0.000
CH4	0.0524	19829.9	1.983
H2O	0.0000	0.432	0.000
Other	0.0002	58.904	0.006
moles	2.64		
%		100.0	

Table 5.C After Methanation Reaction

Assumption: 99.975% Conversion

	moles	ppm	%
H2	51.76682	998766.5	99.877
CO	2.64E-06	0.051	0.000
CO2	1.32E-07	0.003	0.000
CH4	0.0524	1010.4	0.101
H2O	0.0114	220.0	0.022
Other	0.0002	3.001	0.000
moles	51.83		
%		100.0	

Table 5.F After H2 Desorption

Assumption: 150 Times volume ratio

	moles	ppm	%
H2	49.1785	999992.9	99.9993
CO	0.0000	0.000	0.0000
CO2	0.0000	0.000	0.0000
CH4	0.0003	7.099	0.0007
H2O	0.0000	0.000	0.0000
Other	0.0000	0.021	0.0000
moles	49.18		
%		100.0	

66.5% Net Recovery
99.9993 % Purity

FIG. 5