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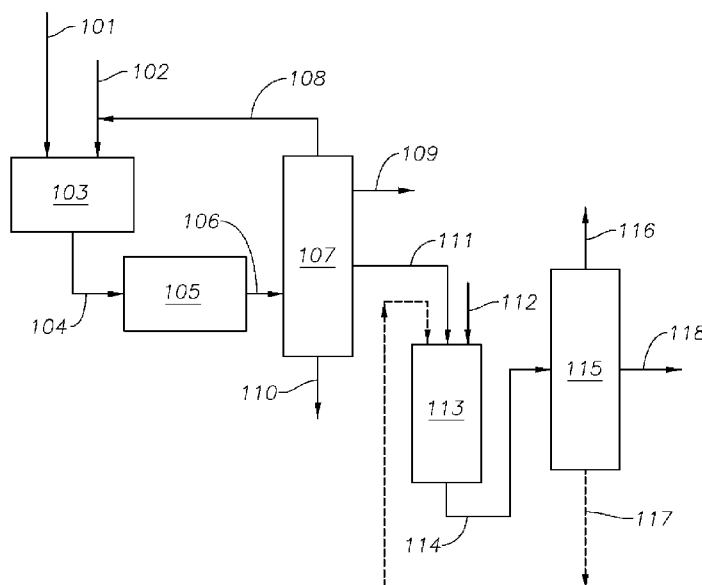
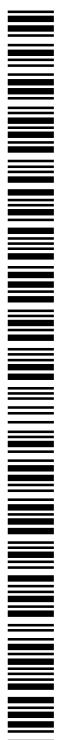


FIG. 1

(57) Abstract: Disclosed herein are processes for producing neopentane. The processes generally relate to demethylating isooctane to produce neopentane. The isooctane may be provided by the alkylation of isobutane with butylenes.



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CROSS REFERENCE TO RELATED APPLICATIONS

[0001] This invention claims priority to and the benefit of USSN 62/380,520, filed August 29, 2016 and EP 16194988.8, filed October 21, 2016.

FIELD OF INVENTION

10 [0002] The present invention relates to methods of producing neopentane and uses thereof.

BACKGROUND OF INVENTION

[0003] Neopentane is a unique nonpolar hydrocarbon molecule that has found industrial use in the form of an inert condensing agent for gas-phase reactions. See, for instance, US 6,262,192. Other potential industrial uses for neopentane include use as a heat removal
15 agent, a blowing agent, and a gasoline blend component due to its relatively high octane numbers. For instance, neopentane has a Research Octane Number (RON) of 85.5 and a Motor Octane Number (MON) of 80.2.

[0004] Currently, there is no satisfactory process for producing neopentane on a commercial scale. For example, typical existing processes for synthesizing neopentane
20 utilize stoichiometric reactions of t-butylchloride and a Grignard reagent, methyl aluminum dichloride, dimethyl aluminum chloride, or trimethyl aluminum. See, for instance, US 3,585,252. Such stoichiometric reactions generate large amounts of metal halides and are difficult to scale up to produce neopentane at commercial quantities. Likewise, though neopentane may be synthesized by hydrogenation of neopentanoic acid under high pressure
25 and at high temperature, e.g., as described in US 4,593,147, such processes are expensive due to the neopentanoic acid feedstock and suffer from a combination of demanding reaction conditions and low selectivity.

[0005] Other proposed processes for producing neopentane involve demethylation of higher carbon number branched paraffins. For example, US 4,940,829 and US 2,422,675
30 each relate to the preparation of neopentane via catalytic demethylation of neohexane. However, these higher carbon number branched paraffins are not readily available in high concentrations suitable as feedstock that could be utilized on a commercial scale.

[0006] Yet alternatively, a process for producing neopentane by hydrogenating an isobutylene polymer and selectively cracking the hydrogenation product is described in US
35 2,394,743. However, in addition to producing neopentane, this process also produces large

amounts of heavier hydrocarbon components.

[0007] Thus, there remains a need for processes for producing neopentane at high yield under mild reaction conditions and utilizing low cost, readily available feedstock. Such processes would allow economic production of neopentane at commercial quantities.

- 5 [0008] Other references of interest include: “The Preparation and Activity for Alkane Reactions of Aerosil-Supported Rhodium-Copper Clusters,” Clarke et al., *Journal of Catalysis*, vol. 111, pp. 374-82 (1988); “Selective Demethylation of Paraffin Hydrocarbons: Preparation of Triptane and Neopentane,” Haensel et al., *Industrial and Engineering Chemistry*, vol. 39, pp. 853-57 (1947); “Skeletal Reactions of Hydrocarbons over Supported
- 10 Iridium-Gold Catalysts,” Fogar et al., *Journal of Catalysis*, vol. 64, pp. 448-63 (1980); “Reactions of 2,2-Dimethylbutane on Iridium: The Role of Surface Carbonaceous Layers and Metal Particle Size,” Vogelzang et al., *Journal of Catalysis*, vol. 111, pp. 77-87 (1988); “Hydrogenolysis of 2,2-Dimethylbutane and n-Hexane over Supported Ruthenium, Nickel, Cobalt, and Iron,” Machiels et al., *Journal of Catalysis*, vol. 58, pp. 268-75 (1979);
- 15 “Hydrogenolysis of Saturated Hydrocarbons: Influence of Hydrocarbon Structures on the Activity and Selectivity of Nickel on Silica,” Leclercq et al., *Journal of Catalysis*, vol. 99, pp. 1-11; GB 574694; US 2,422,670; US 2,436,923; “STRATCO Effluent Refrigerated H₂SO₄ Alkylation Process,” in Handbook of Petroleum Refining Processes, Third Edition, Graves, ch. 1.2 (2004); “UOP Alkylene™ Process for Motor Fuel,” in Handbook of Petroleum
- 20 Refining Processes, Third Edition, Roeseler, ch. 1.3 (2004); and “UOP HF Alkylation Technology,” in Handbook of Petroleum Refining Processes, Third Edition, Himes et al., ch. 1.2 (2004).

SUMMARY OF THE INVENTION

- [0009] The present invention relates to novel processes that address the need for the
- 25 production of neopentane at high yield, under mild reaction conditions, and utilizing readily available feedstock. In one aspect, the present invention relates to a process for producing neopentane comprising alkylating isobutane with butylenes to produce isooctane (2,2,4-trimethylpentane), followed by demethylating the isooctane to produce a product comprising at least 10 wt% neopentane. Typically, the butylenes can be provided in a C₄ olefinic feed
- 30 stream, preferably a refinery raffinate stream, such as a raffinate stream obtained from cracking naphtha. In another aspect the present invention relates to a process for producing neopentane comprising demethylating isooctane to produce a product comprising at least 10 wt% neopentane.

BRIEF DESCRIPTION OF THE DRAWINGS

[0010] FIG. 1 is a diagram of a process of making neopentane.

[0011] FIG. 2 is a graph of conversion of isooctane to neopentane, neohexane, and neoheptane against time on stream (T.O.S.) in the demethylation of Example 7.

5 **DETAILED DESCRIPTION OF THE EMBODIMENTS**

[0012] Described herein are processes for producing neopentane. As discussed below, the processes involve the demethylation of isooctane, preferably via contacting the isooctane with hydrogen in the presence of a catalyst. The isooctane can be provided by the alkylation of isobutane, preferably via contacting the isobutane with butylenes in the presence of a catalyst. Preferably, the butylenes provided in a C₄ olefinic feed, such as a refinery raffinate stream. Alternatively, the isooctane can be provided in an independent feed stream. Preferably, the processes described herein enable the production of neopentane in quantities of greater than about 5 kg/hr, preferably greater than about 500 kg/hr, preferably greater than about 5000 kg/hr, and preferably greater than about 35000 kg/hr.

15 [0013] Unless otherwise indicated, all numbers indicating quantities in the present disclosure are to be understood as being modified by the term “about” in all instances. It should also be understood that the precise numerical values used in the specification and claims constitute specific embodiments. Efforts have been made to ensure the accuracy of the data in the examples. However, it should be understood that any measured data
20 inherently contain a certain level of error due to the limitation of the technique and equipment used for making the measurement.

[0014] As used herein, the indefinite article “a” or “an” shall mean “at least one” unless specified to the contrary or the context clearly indicates otherwise. Thus, embodiments using “a fractionation column” include embodiments where one, two or more fractionation columns
25 are used, unless specified to the contrary or the context clearly indicates that only one fractionation column is used. Likewise, “a C₁₂₊ component” should be interpreted to include one, two or more C₁₂₊ components unless specified or indicated by the context to mean only one specific C₁₂₊ component.

[0015] As used herein, “wt%” means percentage by weight, “vol%” means percentage by
30 volume, “mol%” means percentage by mole, “ppm” means parts per million, and “ppm wt” and “wppm” are used interchangeably to mean parts per million on a weight basis. All “ppm” as used herein are ppm by weight unless specified otherwise. All concentrations herein are expressed on the basis of the total amount of the composition in question. Thus, the concentrations of the various components of the first feedstock are expressed based on the

total weight of the first feedstock. All ranges expressed herein should include both end points as two specific embodiments unless specified or indicated to the contrary.

[0016] Nomenclature of elements and groups thereof used herein are pursuant to the Periodic Table used by the International Union of Pure and Applied Chemistry after 1988.

5 An example of the Periodic Table is shown in the inner page of the front cover of Advanced Inorganic Chemistry, 6th Edition, by F. Albert Cotton et al. (John Wiley & Sons, Inc., 1999).

[0017] As used herein, “hydrocarbon” refers to molecules or segments of molecules containing primarily hydrogen and carbon atoms. As used herein, the term “C_n” hydrocarbon wherein n is a positive integer, e.g., 1, 2, 3, 4, etc., means a hydrocarbon having n number of carbon atom(s) per molecule. The term “C_n+” hydrocarbon wherein n is a positive integer, e.g., 1, 2, 3, 4, etc., as used herein, means a hydrocarbon having at least n number of carbon atom(s) per molecule. The term “C_n-” hydrocarbon wherein n is a positive integer, e.g., 1, 2, 3, 4, etc., used herein, means a hydrocarbon having no more than n number of carbon atom(s) per molecule.

15 [0018] As used herein, “olefin” refers to any unsaturated hydrocarbon having the formula C_nH_{2n} and containing one carbon-carbon double bond, wherein C is a carbon atom, H is a hydrogen atom, and n is the number of carbon atoms in the olefin.

[0019] As used herein, “alkane” or “paraffin” refers to any saturated hydrocarbon having the formula C_nH_{2n+2}, wherein C is a carbon atom, H is a hydrogen atom, and n is the number of carbon atoms in the alkane.

[0020] As used herein, a “primary carbon atom” refers to a carbon atom neighboring one carbon atom, “secondary carbon atom” refers to a carbon atom neighboring two carbon atoms, “tertiary carbon atom” refers to a carbon atom neighboring three carbon atoms, and “quaternary carbon atom” refers to a carbon atom neighboring four carbon atoms.

25 [0021] As used herein, the prefix “normal” or “n-” signifies a linear unbranched hydrocarbon.

[0022] As used herein, the prefix “iso” or “i-” signifies a hydrocarbon containing a methyl substitution at the second carbon of the hydrocarbon chain.

[0023] As used herein, the prefix “neo” signifies a hydrocarbon containing a quaternary carbon atom. For example, the term “neopentane” refers to a compound of the formula C₅H₁₂ and containing a quaternary carbon atom, otherwise known as 2,2-dimethylpropane.

Alkylation of Isobutane

[0024] Often, isooctane is formed in the present invention by the alkylation of isobutane, preferably via contacting the isobutane with butylenes in the presence of a catalyst.

Preferably, the butylenes are provided in a C₄ olefinic feed stream. Suitable C₄ olefinic feeds include C₄ hydrocarbon mixtures obtained in refining, cracking (thermal, catalytic cracking or steam cracking) and/or reforming of oils, butane-butene fractions obtained by removing butadiene from C₄ by-product fractions formed in the production of ethylene by thermal cracking of oils, or C₄ hydrocarbon mixtures obtained by dehydrogenation of hydrocarbon mixtures containing n-butane and isobutane. The C₄ olefinic feed stream preferably comprises a raffinate stream obtained from a refinery or chemical plant cracked naphtha stream, such as from a steam cracker or fluid catalytic cracker.

[0025] Often, the C₄ olefinic feed comprises: from about 5 wt% to about 60 wt% isobutylene, such as from about 10 wt% to 50 wt% or about 20 wt% to 40 wt%; from about 5 wt% to about 50 wt% 1-butene, such as from about 10 wt% to 40 wt%; from about 5 wt% to about 50 wt% n-butane, such as from about 10 wt% to 40 wt% or about 20 wt% to 30 wt%; from about 5 wt% to about 50 wt% cis- and trans-2-butene, such as from about 10 wt% to 40 wt% or about 20 wt% to 30 wt%; and from about 1 wt% to about 20 wt% isobutane, such as from about 5 wt% to 10 wt%, each by weight of the olefinic feed (100 wt%). The C₄ olefinic feed may also have minor amounts (0.01 wt% to 5 wt%) of polar molecules or molecules comprising polar moieties such as nitriles, mercaptans, or oxygenated components. Optionally, the C₄ olefinic feed may further comprise butadiene.

[0026] Preferably, the alkylation is conducted in the presence of a catalyst. The catalyst employed in the alkylation reaction is generally acidic. Any catalyst suitable for isoparaffin alkylation, whether homogeneous or heterogeneous, may be used. Examples of suitable acidic homogeneous catalysts include hydrofluoric acid, sulfuric acid, and mixtures thereof. Examples of suitable acidic heterogeneous catalysts include chlorided alumina, fluorided alumina, zeolites, acidic metal oxides and mixed metal oxides, and mixtures thereof. Non-limiting examples of such zeolites include those of the MOR, BEA, FAU, MTW, and MWW families, preferably the FAU, MWW, and MOR families. Non-limiting examples of acidic metal oxides or mixed metal oxides include tungsten oxides (WO_x), molybdenum oxide (MoO_x), mixed oxides such as WO_x/ZrO₂, WO_x/CeO₂, MoO_x/ZrO₂, MoO_x/CeO₂, and sulfated zirconia.

[0027] The alkylation reaction can be conducted in a wide range of reactor configurations including fixed bed (single or in series), slurry reactors, and/or catalytic distillation towers. In addition, the alkylation reaction can be conducted in a single reaction zone or in a plurality of reaction zones, preferably in a plurality of reaction zones.

[0028] When a homogenous acid catalyst is used, such as hydrofluoric acid or sulfuric acid, suitable reaction temperatures range from about 0°C and about 50°C, such as from about 5°C and about 40°C, or from about 10°C and about 25°C. In such aspects, the molar ratio of isobutane to butylenes supplied to the alkylation reaction zone or zones is typically greater than about 10:1, such as from about 10:1 to about 75:1, or from about 20:1 to about 50:1.

[0029] When a heterogeneous acid catalyst is used, such as a zeolite, suitable reaction temperatures range from about 100°C to about 400°C, such as from about 125°C to about 300°C, or from about 150°C to about 250°C. In such aspects, the molar ratio of isobutane to butylenes supplied to the alkylation reaction zone or zones is typically greater than about 50:1, such as from about 50:1 to about 200:1, or from about 75:1 to about 100:1.

[0030] Irrespective of the catalyst employed, the reaction pressure is preferably maintained so that the C₄ olefinic feed remains in liquid form within the reactor. For instance, suitable reaction pressures are from about 100 kPa absolute to about 7000 kPa absolute (e.g., atmospheric to about 1000 psia), such as from about 500 kPa absolute to about 5000 kPa absolute.

[0031] When a homogeneous acid catalyst is employed, the alkylation reaction effluent is generally subjected to downstream treatment, preferably a water wash step, to remove entrained acid. The water wash step can be conducted in one or more stages and can be carried out in any apparatus known in the art, such as a wash tower. Optionally, the purified alkylation reaction effluent is subsequently dried. Drying may utilize a molecular sieve, such as 3Å molecular sieves. Drying may also be accomplished by pasteurization.

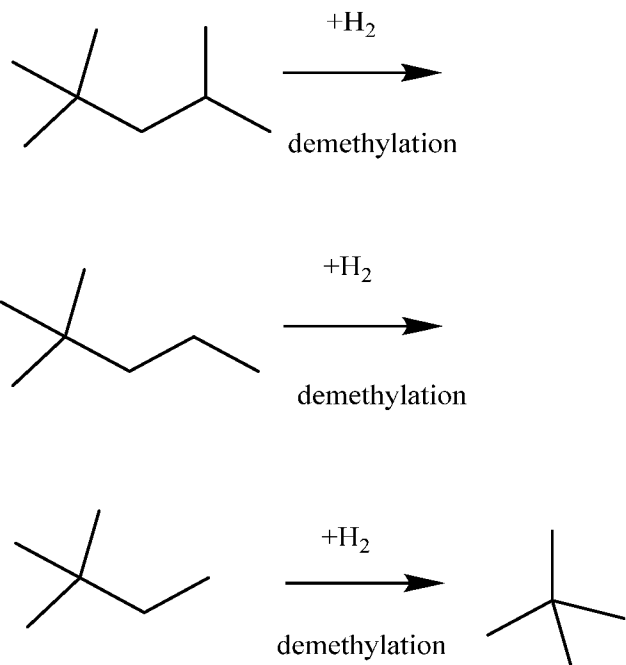
Demethylation of Isooctane

[0032] The major components of the alkylation reaction effluent, whether or not subjected to downstream treatment, are generally unreacted components of the isobutane and C₄ olefinic feeds, primarily n-butane and isobutane, and alkylates. The alkylates typically comprise the desired product, isooctane, and a variety of C₈+ byproducts. The primary constituents of the alkylates, along with each of their boiling points, are summarized in Table 1.

TABLE 1

Alkylate	Boiling Point (°C)
2,2,4-trimethyl pentane (isooctane)	99.2
2,2,3-trimethyl pentane	109.8
2,3,3-trimethyl pentane	114.8
2,3,4-trimethyl pentane	113.5
2,2,3,3-tetramethyl butane	106.5
2-methyl-3-ethyl pentane	115.6
3-methyl-2-ethyl pentane	118.3

- [0033] The unreacted feed components and alkylate byproducts can be readily removed from the reaction effluent by, for example, distillation. The remainder of the alkylation reaction effluent, mainly composed of isooctane, can be demethylated to produce neopentane. Preferably, the separated alkylation reaction effluent comprises greater than about 80 wt% isooctane, or greater than about 90 wt% isooctane, or greater than about 95 wt% isooctane, or greater than about 99 wt% isooctane, such as from about 90 wt% to about 100 wt% isooctane, or from about 95 wt% to about 99 wt% isooctane.
- [0034] Alternatively or additionally, an independent isooctane feed stream can be provided and demethylated. In such aspects, the isooctane feed stream preferably comprises greater than about 80 wt% isooctane, or greater than about 90 wt% isooctane, or greater than about 95 wt% isooctane, or greater than about 99 wt% isooctane, such as from about 80 wt% to about 99 wt% isooctane, or from about 85 wt% to about 95 wt% isooctane.
- [0035] In any embodiment, the isooctane feed stream preferably comprises less than about 30 wt% of aromatic hydrocarbons, such as less than about 10 wt% of aromatic hydrocarbons. Additionally or alternatively, the isooctane feed stream may be free or substantially free of aromatic hydrocarbons.
- [0036] Preferably, the demethylation is conducted by contacting the isooctane with hydrogen in the presence of a catalyst. The reaction pathway for the conversion of the isooctane to neopentane typically proceeds by a step-wise demethylation that may be summarized in the following reaction scheme:



[0037] As shown from the reaction scheme above, the desired demethylation occurs at the tertiary (3°) carbon of the isooctane and the secondary (2°) carbon of the intermediates. Competing demethylation reactions can occur at the quaternary (4°) carbon. Advantageously, demethylation at the quaternary (4°) carbon in the present processes is minimized to prevent a loss of neopentane yield.

[0038] The demethylation reaction can be conducted in a wide range of reactor configurations including fixed bed (single or in series), slurry reactors, and/or catalytic distillation towers. In addition, the demethylation reaction can be conducted in a single reaction zone or in a plurality of reaction zones. The demethylation is conveniently conducted at a temperature from about 200°C to about 500°C, such as from about 300°C to about 400°C and a pressure from about 100 kPa absolute to about 10000 kPa absolute (e.g., atmospheric to about 1500 psia), such as from about 300 kPa absolute to about 8000 kPa absolute, in the presence of a catalyst. Often, the demethylation is conducted at a hydrogen partial pressure of from about 50 kPa absolute to about 3500 kPa absolute (e.g., from about 7 psia to about 500 psia). Preferably, the demethylation is conducted at a hydrogen partial pressure of less than about 2500 kPa absolute, preferably less than about 2200 kPa absolute, and preferably less than about 1000 kPa absolute (e.g., preferably less than about 350 psia, or preferably less than about 150 psia). Particularly preferably, the demethylation may be conveniently conducted under conditions comprising one or more of the following: a temperature from about 220°C to about 300°C; a pressure from about 15 psig to about 200

psig (e.g., from about 205 kPa absolute to about 1400 kPa absolute); and a hydrogen to hydrocarbon molar ratio from about 1:1 to about 14:1.

[0039] Generally, the catalyst employed in the demethylation comprises a transition metal component. Specific, non-limiting examples of suitable transition metal components include
5 Fe, Co, Ni, Rh, Ir, Ru, Pt, and Pd, combinations thereof, compounds thereof, and mixtures of compounds thereof, with Ni being particularly advantageous. Often, the transition metal component contains transition metal as a single component. Alternatively, the transition metal component may contain a transition metal combined with an additional metal to form a binary or ternary alloy. Specific, non-limiting examples of suitable additional metals include
10 Cu, Au, Ag, Sn, Zn, Re, combinations thereof, compounds thereof, and mixtures of compounds thereof. Desirably, the amount of the transition metal component present in the catalyst is from about 0.05 wt% to about 60.0 wt%, such as from about 0.10 wt% to about 50.0 wt%, of the total weight of the catalyst. Generally, the transition metal component is supported on a non-acidic support material. Specific, non-limiting examples of suitable
15 support materials include silica, theta-alumina, clay, pentasil, aluminophosphate, carbon, titania, zirconia, and mixtures thereof.

[0040] Preferably, the acidity of the catalyst employed in the demethylation is minimized to inhibit undesired cracking reactions. Often, the acidity of the catalyst is reduced via impregnation with an alkali metal compound, preferably an alkali metal hydroxide, nitrate,
20 carbonate, bicarbonate, or oxide, such as sodium oxide, e.g., Na₂O. Desirably, the amount of the alkali metal compound present in the catalyst is from about 0.05 wt% to about 1.0 wt%, such as from about 0.1 wt% to about 0.5 wt%, of the total weight of the catalyst.

[0041] Typically, the overall isooctane conversion (i.e., including recycle) during the demethylation step is greater than 50%, preferably greater than 80%, preferably greater than
25 95%, and preferably greater than 99%, such as from 80% to 99% or 90 to 99%. The product of the demethylation step generally comprises neopentane, C₄- hydrocarbon components (e.g., methane, ethane, and propane) and, optionally, partially converted C₆+ hydrocarbon intermediate components (e.g., neohexane and neoheptane). Preferably, the product of the demethylation step comprises: at least 10 wt%, preferably at least about 25 wt%, preferably at
30 least about 35 wt%, and ideally at least about 50 wt% of neopentane, such as from about 25 wt% to about 50 wt% or from about 30 wt% to about 40 wt%; less than about 75 wt%, preferably less than about 65 wt%, and preferably less than about 50 wt% of C₄- hydrocarbon components such as from about 25 wt% to about 75 wt% or from about 40 wt% to about 60 wt%; less than about 5 wt%, preferably less than about 1 wt%, and ideally less than about 0.5

wt% of non-neopentane C₅ hydrocarbon components, such as from about 0 wt% to about 1 wt%; and less than about 10 wt%, preferably less than about 5 wt%, preferably less than about 1 wt%, and ideally less than about 0.5 wt% of partially converted C₆₊ hydrocarbon components (e.g., C₆-C₇ hydrocarbons), such as from about 0 wt% to about 10 wt%, or from
5 about 0 wt% to about 1 wt%, or from about 0.5 wt% to about 1 wt%.

[0042] The light C₄- hydrocarbon components and the C₆₊ hydrocarbon intermediate components can be readily removed from the demethylation product by, for example, distillation, thereby yielding a purified neopentane product stream. Preferably, the purified neopentane product stream comprises greater than about 80 wt% neopentane, or greater than
10 about 90 wt% neopentane, or greater than about 95 wt% neopentane, or greater than about 99 wt% neopentane, such as from about 80 wt% to about 99 wt% neopentane, or from about 85 wt% to about 95 wt% neopentane.

Process

[0043] The present inventive process will now be more particularly described with reference to FIG. 1. FIG. 1 illustrates one aspect of the present inventive process, in which
15 an isobutane stream is alkylated with butylenes provided in a C₄ olefinic feed stream to produce an alkylation product, after which isooctane is separated and demethylated. The invention is not limited to this aspect, and this description is not meant to foreclose other aspects within the broader scope of the invention, such as those where an independent
20 isooctane feed stream is provided and demethylated.

[0044] As shown in FIG. 1, a C₄ olefinic feed stream 101 and an isobutane stream 102 are fed to an alkylation reactor 103 to produce an alkylation effluent 104 comprising isooctane, unreacted isobutane, n-butane, and alkylate byproducts, e.g., C₈₊ hydrocarbons. Optionally, the alkylation effluent is then fed to a treatment unit(s) 105 where the alkylation product is
25 washed with water and subsequently dried to produce a treated effluent 106. The alkylation effluent 105 or treated effluent 106 is then fed to a separator 107, e.g., a distillation column, to separate a first light fraction 108 comprising unreacted isobutane, a second light fraction 109 comprising n-butane, and a heavy fraction 110 comprising alkylate byproducts (e.g., C₈₊ hydrocarbons) from the alkylation effluent or treated effluent. The resulting obtained fraction
30 111 is mainly composed of isooctane. The light fraction 108 can be recycled to alkylation reactor 103. Preferably, the light fraction 109 may be used for fuel (not shown). Preferably, the heavy fraction 110 may be used for mogas (not shown). Fraction 111 and a hydrogen stream 112 are then introduced to a demethylation reactor 113 to produce a demethylation effluent 114 comprising neopentane, C₄- hydrocarbons, and partially converted components,

e.g., C₆+ hydrocarbons. The demethylation effluent 114 is then fed to a separator 115, e.g., a distillation column, to separate a light fraction 116 comprising C₄- hydrocarbons and a heavy fraction 117 comprising partially converted C₆+ hydrocarbons (primarily, C₆-C₇ hydrocarbons) from the demethylation effluent 114. The resulting obtained fraction 118 is
5 mainly composed of neopentane. The light fraction 116 can be subjected to further downstream treatment for use as fuel (not shown). Optionally, the heavy fraction 117 can be recycled to the demethylation reactor.

[0045] Neopentane produced in accordance with the present invention is useful as a blowing agent for the production of foamed polymers and possesses several properties (e.g., a
10 boiling point of 9.5°C and a freezing point of -16.6°C) making it useful as a heat removal agent and/or an inert condensing agent (ICA) in gas phase polymerization process, such as gas phase polymerization processes for the production of polyethylene. Neopentane produced in accordance with this invention also exhibits high octane numbers and is therefore useful as a gasoline blend component.

15 [0046] The invention will now be more particularly described with reference to the accompanying drawings and the following non-limiting Examples.

EXAMPLES

Example 1: Synthesis of 1.5%Ir/SiO₂ Demethylation Catalyst

[0047] A 1.5%Ir/SiO₂ catalyst was prepared by incipient wetness impregnation. Silica gel
20 (Davisil™ 646, Sigma-Aldrich) was calcined by heating at a rate of 10°C/min to 700°C and holding at 700°C for 15 hours, then cooled to 50°C and held at temperature overnight. The pore volume of the calcined silica was determined to be 1.34 cc/g. An aqueous solution of iridium (III) chloride was prepared by adding 0.28 g of iridium (III) chloride hydrate to 13.4 mL water, followed by adding 5 drops of concentrated HCl. The resulting mixture was then
25 stirred and heated until the metal salt completely dissolved. The calcined silica (10 g) was then impregnated with the prepared aqueous solution to give an Ir loading of 1.5 wt%. The resultant product was transferred to a ceramic dish, calcined by heating at a rate of 10°C/min to 250°C and holding at 250°C for 10 hours, then cooled to 50°C and held at temperature overnight.

Example 2: Synthesis of 1.5%Rh/SiO₂ Demethylation Catalyst

30 [0048] A 1.5%Rh/SiO₂ catalyst was prepared by incipient wetness impregnation. Silica gel (Davisil™ 646, Sigma-Aldrich) was calcined by heating at a rate of 10°C/min to 700°C and holding at 700°C for 15 hours, then cooled to 50°C and held at temperature overnight. The pore volume of the calcined silica was determined to be 1.34 cc/g. The calcined silica (10 g)

was then impregnated with a solution of rhodium (III) chloride hydrate (0.39 g) dissolved in 13.4 mL water to give an Rh loading of 1.5 wt%. The resultant product was transferred to a ceramic dish, calcined by heating at a rate of 10°C/min to 250°C and holding at 250°C for 10 hours, then cooled to 50°C and held at temperature overnight.

5 Example 3: Synthesis of Na₂O/Rh/SiO₂ Demethylation Catalysts

[0049] Na₂O/Rh/SiO₂ catalysts with varying Na₂O content were prepared from the calcined 1.5%Rh/SiO₂ catalyst of Example 1. The 1.5%Rh/SiO₂ catalyst was first impregnated with an aqueous solution of NaNO₃ to achieve the desired Na₂O loading (e.g., 0.1 w% or 0.5 wt%). The Na₂O/Rh/SiO₂ catalyst was then obtained by drying the impregnated sample
10 overnight at 100°C followed by heating to 400°C under flowing nitrogen.

Demethylation of Isooctane and GC Characterization Procedures for Examples 4-7

[0050] All demethylation tests described in Examples 4-7 were carried out in a down-flow, tubular, fixed bed reactor equipped with two 100-mL ISCO pumps and various gas feeds. The liquid feed was delivered via the ISCO pumps, mixed with gas feed through a heated
15 section for vaporization before entering the reactor (3/8 in O.D. × 16 3/4 in × 0.028 in wall stainless steel tube) (1 cm × 43 cm × .07 cm). Catalyst (0.25 – 2 g loading) was pelletized and sized to 20-40 mesh, diluted with quartz chips to a total volume of 5 mL and loaded in the isothermal zone of the reactor. A piece of 1/4 in (0.6 cm) O.D. stainless steel tubing was inserted at the bottom of the reactor tube to ensure the catalyst bed being located in the
20 isothermal zone of the furnace. Glass wool was used at the top and bottom of the catalyst bed to keep the catalyst bed in place. Reactor pressure was controlled via a research control valve (RCV) at the exit of the reactor and the reactor effluent was heat-traced and sent to a gas chromatograph (GC) for on-line analysis.

[0051] The catalyst was first purged with N₂ and then heated to 300 – 500°C at a ramp rate
25 of 3°C/min under flowing H₂ (100 cc/min) and held 2–4 h for reduction. After reduction, the reactor was cooled down to the operating temperature. The liquid feed was then introduced and H₂ flow rate adjusted accordingly at the desired operating pressure.

[0052] An Agilent HP 7890 GC having a Restek 30m × 0.32mm × 5 μm GC column (Rtx™-1, Catalog # 10178) was used for product analysis in Examples 4-7. The injector was
30 set at 260°C and the detector at 280°C. The column flow rate was 1.2 cc/min He, with typical air and H₂ flows for the detector. The oven temperature was programmed in the following manner: initial temperature of 40°C; hold for 5 min; ramp at 4°C/min to 200°C; ramp at 20°C/min to 260°C for bake out. The total reactor effluent was sampled and analyzed hourly. GC peaks were identified using the “Alphagaz PIANO Calibration

Standards” (Supelco Product # 4-4586-U, available from Sigma-Aldrich) and authentic samples, using a response factor of one for all components.

Example 4: Demethylation of Isooctane Using 1.5%Ir/SiO₂ Catalyst

[0053] The 1.5%Ir/SiO₂ catalyst of Example 1 (1.019 g) was sized to 20/40 sieve mesh, dispersed with quartz chips (20/40 mesh) then loaded into the isothermal zone of the reactor to a volume of 5.5 cc. The catalyst was pre-conditioned *in situ* by heating to 400°C with H₂ flow at 100 cc/min and holding for 2 h. The feed was mixed with hydrogen and pumped through heated lines for vaporization prior to flowing to the reactor. The feed was then pumped through the catalyst bed at initial reaction conditions comprising a temperature of 300°C at a liquid weight hourly space velocity (LWHSV) of 4 h⁻¹, a hydrogen flow rate of 142 cc/min, and a pressure of 15 psia (100 kPa absolute).

[0054] GC testing of the reactor effluent at a temperature of 300°C and a pressure of 15 psia yielded the following results: isooctane conversion of 30%; neoheptane selectivity of 4.8 wt%; neohexane selectivity of 4.8%; and a neopentane selectivity of 1.9 wt%.

Example 5: Demethylation of Isooctane Using Rh/SiO₂ Catalyst with varying Na₂O content

[0055] Catalyst prepared in accordance with Example 4 (0.25-2 g, e.g., 1.049 g 0%Na₂O/1.5%Rh/SiO₂) was sized to 20/40 sieve mesh, dispersed with quartz chips (20/40 mesh) then loaded into the isothermal zone of the reactor to a volume of 5.5 cc. The catalyst was pre-conditioned *in situ* by heating to 400°C with H₂ flow at 100 cc/min and holding for 2 h. The feed was mixed with hydrogen and pumped through heated lines for vaporization prior to flowing to the reactor. The feed was then pumped through the catalyst bed at initial reaction conditions comprising a temperature of 350°C at a LWHSV of 4 h⁻¹, a hydrogen flow rate of 142 cc/min, and a pressure of 15 psia (100 kPa absolute).

[0056] The results of the GC testing of the reactor effluent at a temperature of 350°C, a pressure of 15 psia, and varying Na₂O content are summarized in Table 2.

TABLE 2

Catalyst	Conversion and Selectivity												
Na ₂ O Content (wt%)	Isooctane Conversion (%)	Methane – (wt%)	Ethane – (wt%)	Propane – (wt%)	C4 – (wt%)	Neo-pentane – (wt%)	Other C5 – (wt%)	Neo-hexane – (wt%)	Other C6 – (wt%)	Neo-heptane – (wt%)	Other C7 – (wt%)	Other C8 – (wt%)	Other
0	19.66	55.02	3.98	2.32	4.18	6.90	0.71	13.33	0.15	7.25	1.72	2.64	1.78
0.1	19.93	59.60	3.64	2.13	3.77	7.99	0.58	11.76	0.12	5.77	1.43	1.91	1.32
0.5	19.89	64.56	2.82	1.80	3.68	9.01	0.47	9.20	0.07	4.42	1.07	1.65	1.25

[0057] Table 2 illustrates that the selectivity towards neopentane increased with increasing Na₂O content. This result was likely attributable to the titration effect of the Na₂O on residual acidity of the Rh/SiO₂ catalyst.

Example 6: Demethylation of Isooctane Using Rh/SiO₂ Catalyst with varying Reaction

5 Pressure

[0058] The demethylation kinetics of isooctane at a variety of reaction pressures was measured by comparing the rate of the desired demethylation of the tertiary carbon, $r(3^{\circ}\text{C})$, to the rate of the undesired demethylation of the quaternary carbon, $r(4^{\circ}\text{C})$ [$r(3^{\circ}\text{C})/r(4^{\circ}\text{C})$] in accordance with the following procedure.

10 [0059] The Rh/SiO₂ catalyst of Example 2 (0 wt% Na₂O content) was loaded and pre-conditioned in accordance with Example 5. At each of the reaction pressures tested, the isooctane feed was pumped through the catalyst bed at a variety of flow rates in order to vary the residence time of the feed within the reactor. The $r(3^{\circ}\text{C})/r(4^{\circ}\text{C})$ ratio was obtained by comparing the selectivity of the primary demethylation product at the varied residence times
15 and extrapolating these results to a zero residence time. The results of the kinetics measurements are summarized in Table 3.

TABLE 3

Pressure (psia)	Pressure (kPa absolute)	$r(3^{\circ}\text{C})/r(4^{\circ}\text{C})$
30	200	8.5
115	800	2.5
315	2000	1.3

[0060] Table 3 illustrates that the $r(3^{\circ}\text{C})/r(4^{\circ}\text{C})$ ratio decreased with increasing reactor
20 pressure. This result indicates that a low reaction pressure is desirable to obtain the desired neopentane product.

Example 7: Demethylation of Isooctane Using Ni/Kieselguhr Catalyst with varying Isooctane Flow Rate

[0061] A Ni/Kieselguhr catalyst (60% Ni, Sigma-Aldrich) (2.066 g) was sized to 20/40
25 sieve mesh, dispersed with quartz chips (20/40 mesh) then loaded into the isothermal zone of the reactor to a volume of 5.5 cc. The catalyst was pre-conditioned *in situ* by heating to 500°C with H₂ flow at 100 cc/min and holding for 4 h. The feed was mixed with hydrogen and pumped through heated lines for vaporization prior to flowing to the reactor. The feed was then pumped through the catalyst bed at initial reaction conditions comprising a
30 temperature of 250°C at LHSV of 1 h⁻¹ or 2 h⁻¹, a hydrogen flow rate of 71 cc/min, and a pressure of 15 psia (100 kPa absolute).

[0062] Typical results of the GC testing of the reactor effluent at a temperature of 250°C and a pressure of 15 psia are summarized in Table 4. In addition, the conversion of isooctane and selectivity to neopentane, neohexane, and neoheptane against T.O.S. at the two tested isooctane flow rates are shown in Fig. 2.

5

TABLE 4

Isooctane											
Flow Rate	Conversion and Selectivity										
LWHSV (h ⁻¹)	Isooctane Conversion (%)	Methane – (wt%)	Ethane – (wt%)	Propane – (wt%)	C4 – (wt%)	Neo- pentane- (wt%)	Other C5 – (wt%)	Neohexane – (wt%)	Other C6 – (wt%)	Neo- heptane – (wt%)	Other C7 – (wt%)
1	100	47.7	3.5	4.3	2.3	40.4	0	1.8	0	0	0
2	99.9	30.7	1.0	1.1	3.3	15.9	1.9	37.4	0.3	7.4	0.9

[0063] Table 4 and FIG. 2 illustrate that the selectivity towards neopentane decreased with increasing isooctane flow rate. Conversely, the selectivity towards neohexane and other C₆+ hydrocarbons increased with increasing isooctane flow rate. These results indicate that at high isooctane flow rate (i.e., short residence time) the step-wise demethylation of isooctane to neopentane did not run to completion and intermediate products were obtained.

Example 8: Demethylation of Isooctane Using Ni/SiO₂ Catalyst with varying Process Conditions

[0064] The demethylation tests described in Examples 8A-8M were carried out in a reactor unit having eight modules serviced by communal gas and liquid feed lines. Each individual module featured independent temperature, pressure, and feed flow controls.

[0065] A Ni/SiO₂ catalyst (64 wt% Ni powder on silica, obtained from Strem Chemicals, Inc.) was pelletized, crushed, sieved (40–20 mesh, 840 to 400 μm), and loaded into each of the unit modules along with SiC diluent. The catalyst loading amount was varied depending on the desired WHSV of the isooctane feed. The catalyst was pre-conditioned *in situ* by heating to 400°C with H₂ flow at 500 sccm/min at 30 psig (300 kPa absolute) and holding for 8 h.

[0066] An Agilent HP 7890 GC equipped with dual inlets (i.e., a front inlet and a rear inlet), two FID detectors, and two 60 m × 250 μm × 1.0 μm HP-1 GC columns (Agilent Technologies) located in parallel (corresponding to the front inlet and rear inlet, respectively) was used for online product analysis to determine the reported conversion and selectivity values. Modules 1-4 of the reactor unit were connected in series with the front inlet, and modules 5-8 of the reactor unit were connected to the rear inlet. Sample injection was enabled via a 250 μl sample loop that was connected to a Valco 6 port GC injection valve. In

order to increase the sampling frequency, two simultaneous injections were done for both inlets.

[0067] The GC was operated in a ramped pressure and split mode using hydrogen as carrier gas and a split ratio of 100 to 1. The initial pressure was set at 20 psi (140 kPa) and held for 1.5 min, and then ramped at 7 psi/min (50 kPa/min) to a final pressure of 50 psi (340 kPa). The initial oven temperature was set at 35°C and held for 2 min, and then ramped at 25°C/min to 250°C. The total analysis time was about 10.6 min, which enabled an injection every 15 min and a corresponding analysis frequency of 1h per pair of modules.

[0068] In each of examples 8A-8M, the WHSV of the isooctane feed was varied in order to vary the residence time of the feed within the reactor. Results of the GC testing at varying process conditions (temperature, total pressure, hydrogen to hydrocarbon (H₂:HC) molar ratio, and WHSV) are summarized in Table 5.

TABLE 5

	Process Conditions				Conversion and Selectivity						
	Temperature (°C)	Total pressure (psig)	H ₂ :HC (mol:mol)	WHSV (h ⁻¹)	Isocutane Conversion (%)	Neohexane Selectivity (%mol)	2,4-dimethylpentane Selectivity (%mol)	Neohexane Selectivity (%mol)	2-methylpentane Selectivity (%mol)	Neopentane Selectivity (%mol)	Isopentane Selectivity (%mol)
Example											
8A	220	150	8	0.45	22.25	29.50	6.04	6.68	1.18	0.22	0.16
				1.1	4.66	37.68	7.17	2.54	0.30	0.00	0.00
				2	8.63	35.61	6.81	3.54	0.50	0.07	0.00
				12	0.77	46.93	7.66	2.75	0.00	0.00	0.00
8B				0.5	61.23	18.57	2.22	12.54	1.80	1.37	0.84
				0.75	50.33	22.84	2.94	10.96	1.74	0.73	0.48
				1	58.58	19.66	2.94	11.84	2.02	1.09	0.73
	230	100	14	1.5	38.24	26.61	4.01	8.65	1.54	0.45	0.29
				2	29.35	29.32	4.43	7.17	1.19	0.31	0.19
				5	11.45	34.84	5.31	4.55	0.60	0.13	0.00
8C	235	32	5	2	51.00	20.03	0.53	13.52	0.59	1.80	0.43
				4	79.68	14.83	0.30	15.91	0.69	2.49	0.63

TABLE 5 (cont.)

Process Conditions		Conversion and Selectivity									
8D	235 100 14	0.5	64.75	18.07	1.91	13.06	1.75	1.60	0.86		
		1	73.90	14.99	1.93	14.06	2.20	1.61	1.05		
		2	44.03	27.30	3.78	10.40	1.58	0.59	0.40		
		3	31.27	29.16	4.06	7.51	1.27	0.34	0.21		
		4	41.16	27.51	3.15	9.74	1.12	1.46	0.66		
8E	240 32 5	5	27.52	30.05	3.80	7.38	1.16	0.33	0.19		
		2	68.08	15.05	0.27	16.13	0.48	2.75	0.53		
8F	240 100 3	4	82.45	13.67	0.22	16.25	0.51	3.13	0.57		
		1	28.14	29.78	1.74	9.86	0.49	0.48	0.16		
		2	31.08	28.64	1.82	10.11	0.61	0.53	0.18		
		3	26.90	31.27	2.02	8.83	0.58	0.43	0.14		
		4	11.03	30.25	1.84	5.66	0.33	0.28	0.13		
		6	13.03	35.57	2.41	6.39	0.35	0.29	0.06		
8G	240 100 4	15	23.42	31.10	2.80	7.56	0.85	0.37	0.17		
		1	54.98	21.85	1.25	13.62	0.85	1.38	0.50		
		3	47.47	25.13	2.12	10.85	1.35	0.74	0.42		
		6	18.23	34.31	2.42	6.98	0.44	0.30	0.08		
8H	240 100 5	15	19.29	32.48	3.10	6.85	0.74	0.31	0.13		
		1.2	44.20	25.02	1.79	11.98	0.88	0.67	0.29		
		1.22	69.68	17.48	1.14	15.49	1.22	1.32	0.61		
		2.7	31.45	28.10	2.24	7.97	0.66	0.37	0.15		
		4	72.63	17.43	1.38	14.10	1.78	1.46	0.82		
		5	44.05	26.22	2.24	10.12	1.25	0.69	0.38		
		30	8.12	35.87	3.77	4.98	0.40	0.19	0.05		

TABLE 5 (cont.)

Process Conditions		Conversion and Selectivity									
8I	240 100 14	0.5	92.91	4.28	0.35	13.28	0.72	6.58	1.83		
		1	68.35	17.34	2.04	12.86	2.01	1.61	0.94		
		1.5	77.74	14.57	1.60	14.36	2.20	1.70	1.08		
		2	65.59	19.17	2.21	12.54	2.06	1.08	0.72		
		3	47.15	25.11	3.10	9.70	1.57	0.55	0.38		
		5	30.57	29.11	3.76	7.70	1.28	0.36	0.22		
		7.5	14.32	40.94	4.74	6.68	0.76	0.24	0.08		
8J	240 150 8	0.45	99.83	0.37	0.03	8.80	0.30	10.04	2.05		
		1.1	59.08	17.49	2.70	10.77	1.35	2.71	1.05		
		2	48.53	22.31	3.45	10.73	1.72	0.93	0.61		
		12	5.38	47.77	6.76	4.52	0.39	0.10	0.00		
8K	247 100 14	3	79.34	14.90	1.30	14.38	2.09	1.74	1.05		
		4	67.44	19.18	1.82	12.77	1.98	1.17	0.73		
		6	50.85	24.42	2.52	10.36	1.68	0.69	0.43		
8L	255 100 14	0.5	57.93	21.05	1.53	11.19	1.34	1.59	0.77		
		1.5	78.04	16.08	1.22	13.39	2.10	1.45	1.05		
		2	66.89	20.23	1.72	11.72	2.04	0.96	0.74		
		3	50.14	25.33	2.46	9.31	1.76	0.55	0.44		
		5	77.34	14.58	0.97	14.17	1.69	2.25	1.06		
		10	51.63	23.20	1.96	10.92	1.59	0.98	0.53		
		15	36.42	27.35	2.48	9.03	1.29	0.65	0.33		
		20	29.54	29.15	2.73	8.11	1.13	0.55	0.25		
8M	267 100 14	15	77.09	14.97	0.73	13.81	1.47	2.49	0.98		
		20	66.92	18.43	1.01	12.85	1.49	1.85	0.79		
		30	47.47	24.00	1.61	10.64	1.40	1.09	0.50		

- [0069] From the data obtained in Examples 8A-8M, the absolute rates of the reactions producing the primary demethylation products and the relative rates of these reactions compared to others were regressed by comparing the selectivity of the demethylation products at the varied residence times and extrapolating these results to a zero residence time.
- 5 The regressed absolute and relative reaction rates were then used to calculate simulated neopentane yields and production rates at a conversion of 80% in a theoretical demethylation process comprising demethylating an isooctane feed and recycling the produced C₆+ hydrocarbon fraction. The obtained simulated neopentane yields and ratios of the simulated neopentane production rates relative to the simulated neopentane production rate of Example
- 10 8A at varying process conditions (temperature, total pressure, and hydrogen to hydrocarbon (H₂:HC) molar ratio) are summarized in Table 6.

TABLE 6

Example	Temperature (°C)	Pressure (psig)	H ₂ :HC (mol:mol)	Simulated Neopentane Yield at 80% Conversion	Ratio of Simulated Neopentane Production Rate Relative to Simulated Neopentane Rate of Example 8A at 80% Conversion
8A	220	150	8	47	1.0
8B	230	100	14	49	2.4
8C	235	32	5	53	14.7
8D	235	100	14	49	3.3
8E	240	32	5	55	20.0
8F	240	100	3	52	9.9
8G	240	100	4	52	8.0
8H	240	100	5	50	7.0
8I	240	100	14	49	4.5
8J	240	150	8	47	3.5
8K	247	100	14	50	6.9
8L	255	100	14	51	11.0
8M	267	100	14	51	21.5

- 15 [0070] As can be seen from Table 6, the simulated neopentane production rates varied significantly with changing process conditions, as expected. However, despite the wide variance in the simulated neopentane production rates, the simulated neopentane yield at 80% remained surprisingly near constant with respect to variances in each of temperature,

pressure, and H₂:HC ratio. These simulated results suggest that the demethylation step is surprisingly robust over the preferred process conditions comprising a temperature from about 220°C to about 300°C; a pressure from about 15 psig to about 200 psig (e.g., from about 205 kPa absolute to about 1400 kPa absolute); and/or a hydrogen to hydrocarbon molar ratio from about 1:1 to about 14:1.

[0071] All documents described herein are incorporated by reference herein, including any priority documents and/or testing procedures to the extent they are not inconsistent with this text. As is apparent from the foregoing general description and the specific embodiments, while forms of the invention have been illustrated and described, various modifications can be made without departing from the spirit and scope of the invention. Accordingly, it is not intended that the invention be limited thereby. Likewise, the term “comprising” is considered synonymous with the term “including.” Likewise, whenever a composition, an element or a group of elements is preceded with the transitional phrase “comprising,” it is understood that we also contemplate the same composition or group of elements with transitional phrases “consisting essentially of,” “consisting of,” “selected from the group of consisting of,” or “is” preceding the recitation of the composition, element, or elements and vice versa.

CLAIMS:

What is claimed is:

1. A process for producing neopentane, the process comprising:
 - (a) contacting isobutane with butylenes under alkylation conditions effective to produce
5 an alkylation product including isooctane; and
 - (b) demethylating the isooctane to produce a demethylation product including at least
10 10 wt% neopentane based on the weight of the demethylation product.
2. The process of claim 1, wherein at least part of the isooctane is separated from the
10 alkylation product prior to demethylation.
3. The process of claim 2, wherein separating the isooctane from the alkylation product
comprises distillation.
- 15 4. The process of claim 3, wherein the alkylation product is separated into fractions
comprising (1) an unreacted isobutane fraction, (2) an isooctane fraction, and (3) a non-
isooctane C₈+ hydrocarbon fraction.
5. The process of claim 4, wherein the contacting is carried out in a reaction vessel, and
20 wherein the unreacted isobutane fraction is recycled to the reaction vessel.
6. The process of any of claims 1 to 5, wherein the butylenes are provided in a raffinate
stream obtained from cracking a naphtha stream.
- 25 7. The process of any of claims 1 to 6, wherein the contacting is carried out in the
presence of a catalyst (A) comprising sulfuric acid, hydrofluoric acid, or mixtures
thereof.
8. The process of claim 7, further comprising washing the alkylation product with water.
30
9. The process of claim 7 or 8, wherein the contacting is carried out at a temperature
ranging from about 0°C to about 50°C and wherein the molar ratio of isobutane to
butylenes supplied to the contacting ranges from about 10:1 to about 75:1.

10. The process of any of claims 1 to 6, wherein the contacting is carried out in the presence of a catalyst (B) selected from the group consisting of chlorided alumina, fluorided alumina, acidic metal oxide, acidic mixed metal oxide, zeolites, and mixtures thereof.
- 5
11. The process of claim 10, wherein the contacting is carried out at a temperature ranging from about 100°C to about 400°C and wherein the molar ratio of isobutane to butylenes supplied to the contacting ranges from about 50:1 to about 200:1.
- 10 12. A process for producing neopentane, the process comprising:
- (a) providing a stream including isooctane; and
- (b) demethylating the isooctane to produce a demethylation product including at least 10 wt% neopentane based on the weight of the demethylation product.
- 15 13. The process of any of claims 1 to 12, further comprising separating at least part of the neopentane from the demethylation product.
14. The process of claim 13, wherein separating the neopentane from the demethylation product comprises distillation
- 20
15. The process of claim 14, wherein the demethylation product is separated into fractions comprising (1) a C₄- hydrocarbon fraction, (2) a neopentane fraction, and (3) a C₆+ hydrocarbon fraction.
- 25 16. The process of claim 15, wherein the demethylation is carried out in a reaction vessel, and wherein the C₆+ hydrocarbon fraction is recycled to the reaction vessel.
17. The process of any of claims 1 to 16, wherein the demethylation comprises contacting the isooctane with hydrogen in the presence of a catalyst (C) comprising at least one
- 30 member selected from the group consisting of Fe, Co, Ni, Rh, Ir, Ru, Pt, Pd, combinations thereof, compounds thereof, and mixtures of compounds thereof.

18. The process of claim 17, wherein the catalyst (C) further comprises at least one member selected from the group consisting of Cu, Au, Ag, Sn, Zn, Re, combinations thereof, compounds thereof, and mixtures of compounds thereof.
- 5 19. The process of claim 17 or 18, wherein the catalyst (C) comprises a support material selected from the group consisting of silica, theta-alumina, clay, pentasil, aluminophosphate, carbon, titania, zirconia, and mixtures thereof.
- 10 20. The process of any of claims 1 to 19, wherein the demethylation is carried out at a temperature of about 200°C to about 500°C and a hydrogen partial pressure of from about 50 kPa absolute to about 3500 kPa absolute (about 7 psia to about 500 psia).
- 15 21. The process of any of claims 1 to 19, wherein the demethylation is carried out under conditions comprising at least one of a temperature from about 220°C to about 300°C, a pressure from about 15 psig to about 200 psig (about 205 kPa absolute to about 1400 kPa absolute), or a hydrogen to hydrocarbon molar ratio from about 1:1 to about 14:1.
- 20 22. The process of any of claims 1 to 21 wherein the demethylation product comprises from about 25 wt% to about 50 wt% neopentane based on the weight of the demethylation product and from about 0 wt% to about 10 wt% C₆+ hydrocarbon components based on the weight of the demethylation product.
- 25 23. A composition comprising neopentane produced in accordance with any of claims 1 to 22.
- 30 24. An article comprising the composition of claim 23, wherein the article is selected from the group consisting of a gasoline blend, a blowing agent, a heat removal agent, and mixtures or combinations thereof.

1/2

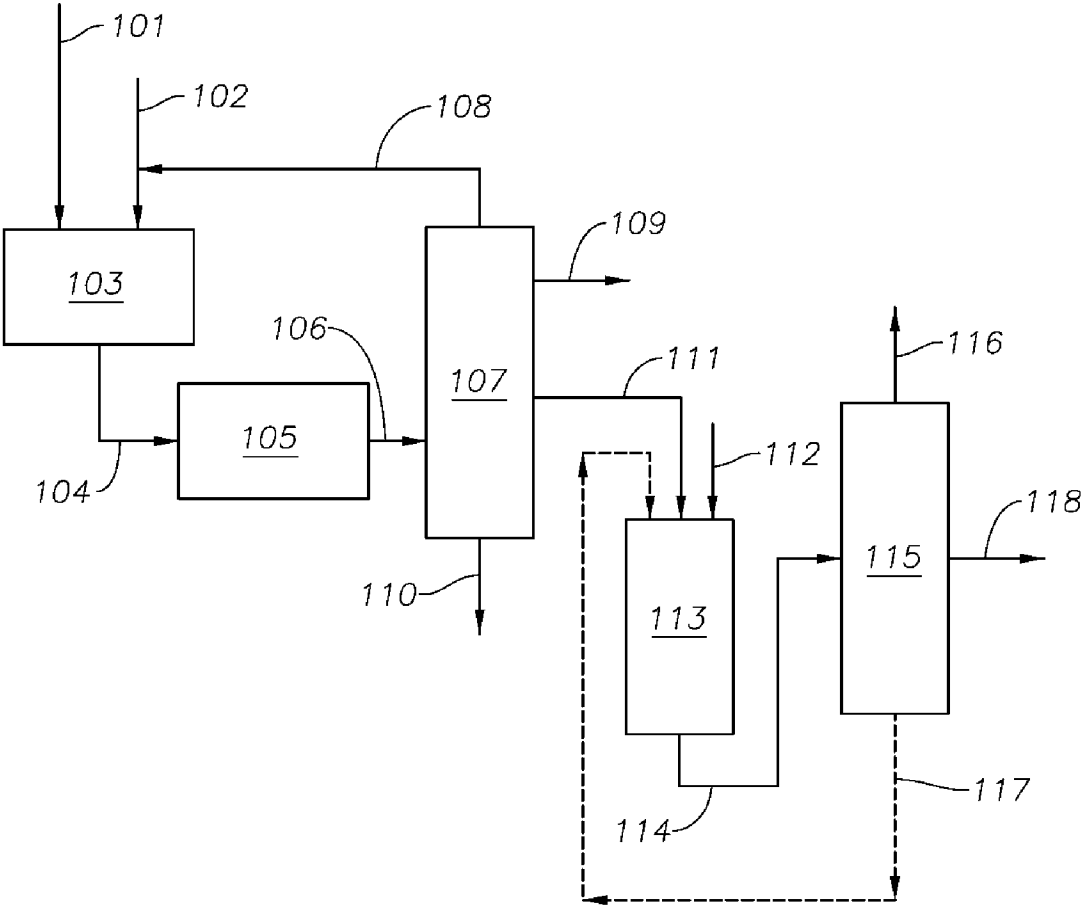


FIG. 1

Ni/Kieselguhr: 15 psia , 250 C, H₂/HC = 14

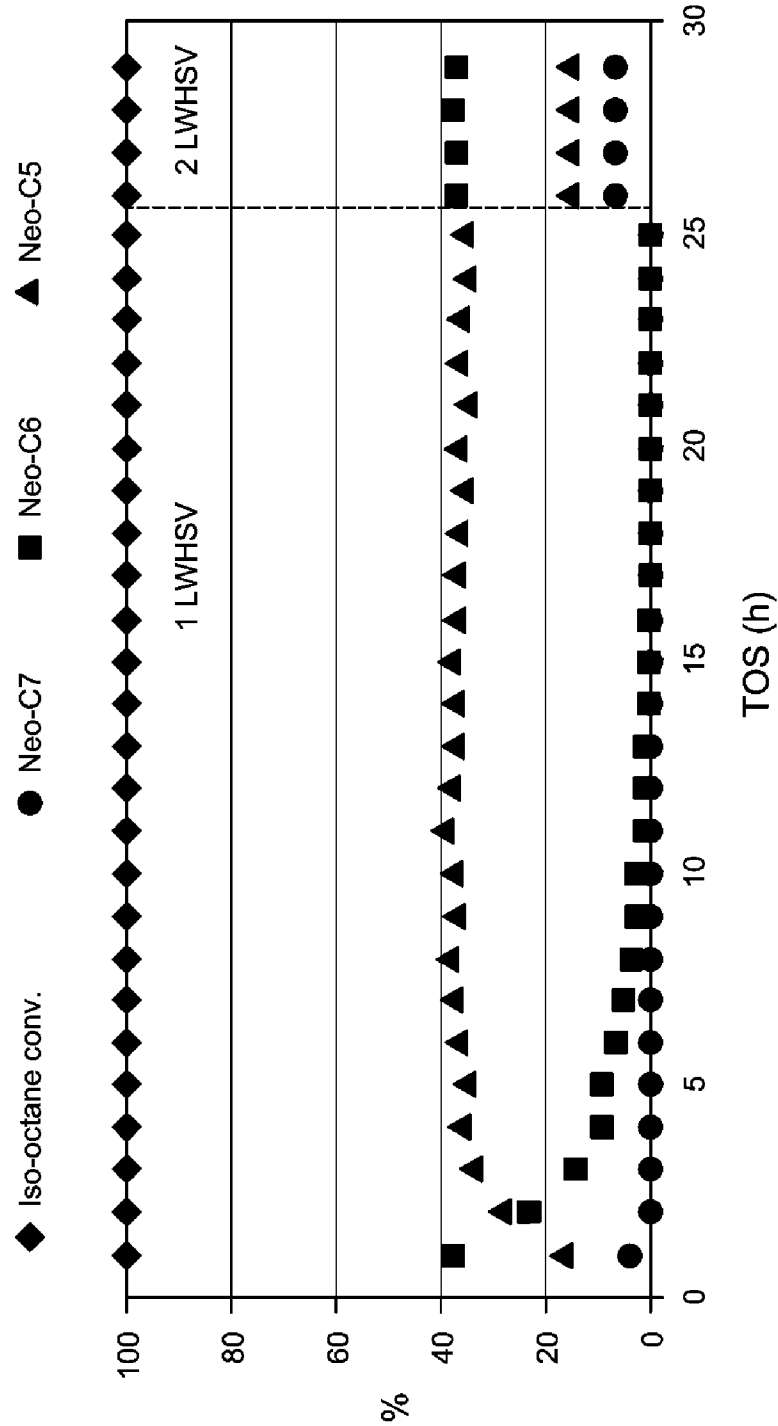


FIG. 2

INTERNATIONAL SEARCH REPORT

International application No.
PCT/US2017/047596**A. CLASSIFICATION OF SUBJECT MATTER****C07C 2/62(2006.01)i, C07C 4/10(2006.01)i, C07C 7/04(2006.01)i, C07C 9/18(2006.01)i, B01J 8/00(2006.01)i**

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

C07C 2/62; C07C 4/10; C07C 7/04; C07C 9/18; B01J 8/00

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Korean utility models and applications for utility models

Japanese utility models and applications for utility models

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

eKOMPASS(KIPO internal), STN (Registry, Caplus), Google & Keywords: neopentane (2,2-dimethylpropane), isooctane (2,2,4-trimethylpentane), alkylation, demethylation

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 2436923 A (HAENSEL, V.) 2 March 1948	12
Y	See claim 7; column 4, lines 68-74; and column 5, lines 46-70.	1-6
Y	KRANZ, K., "Alkylation chemistry-Mechanisms, operating variables, and olefin interactions", DuPont Company, 2003 See page 4.	1-6
Y	US 2394743 A (BERGSTEINSSON, I.) 12 February 1946 See page 1, left column, lines 3-4; page 2, left column, line 72 - right column, line 14; and example.	1-6
A	US 2422672 A (HAENSEL, V. et al.) 24 June 1947 See claim 1; and column 2, lines 21-23.	1-6, 12
A	ZIMMER, H. et al., "Hydrogenolysis of alkanes with quaternary carbon atoms over Pt and Ni black catalysts", J. Chem. Soc., Faraday Trans. 1, 1982, Vol. 78, pages 3573-3586 See page 3573; and table 3.	1-6, 12



Further documents are listed in the continuation of Box C.



See patent family annex.

* Special categories of cited documents:

"A" document defining the general state of the art which is not considered to be of particular relevance

"E" earlier application or patent but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

"&" document member of the same patent family

Date of the actual completion of the international search

20 November 2017 (20.11.2017)

Date of mailing of the international search report

20 November 2017 (20.11.2017)

Name and mailing address of the ISA/KR

International Application Division

Korean Intellectual Property Office

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INTERNATIONAL SEARCH REPORTInternational application No.
PCT/US2017/047596**Box No. II Observations where certain claims were found unsearchable (Continuation of item 2 of first sheet)**

This international search report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☐ Claims Nos.:
because they relate to subject matter not required to be searched by this Authority, namely:
2. ☒ Claims Nos.: 8, 11, 14, 15, 16, 18, 24
because they relate to parts of the international application that do not comply with the prescribed requirements to such an extent that no meaningful international search can be carried out, specifically:
Claims 8, 11, 14, 15, 16, 18 and 24 refer to claims which are not drafted in accordance with the third sentence of Rule 6.4(a).
3. ☒ Claims Nos.: 7, 9, 10, 13, 17, 19-23
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box No. III Observations where unity of invention is lacking (Continuation of item 3 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

1. ☐ As all required additional search fees were timely paid by the applicant, this international search report covers all searchable claims.
2. ☐ As all searchable claims could be searched without effort justifying an additional fees, this Authority did not invite payment of any additional fees.
3. ☐ As only some of the required additional search fees were timely paid by the applicant, this international search report covers only those claims for which fees were paid, specifically claims Nos.:
4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this international search report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest and, where applicable, the payment of a protest fee.
- ☐ The additional search fees were accompanied by the applicant's protest but the applicable protest fee was not paid within the time limit specified in the invitation.
- ☐ No protest accompanied the payment of additional search fees.

INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No.

PCT/US2017/047596Patent document
cited in search reportPublication
datePatent family
member(s)Publication
date

US 2436923 A

02/03/1948

None

US 2394743 A

12/02/1946

None

US 2422672 A

24/06/1947

None