

3.40	25.99	100
6.04	14.62	3
6.81	12.97	81
9.03	9.78	29
11.84	7.47	20
13.01	6.80	3
13.69	6.46	4
15.70	5.64	10
16.10	5.50	31
16.59	5.34	8
17.17	5.16	4
18.13	4.89	14
20.84	4.26	23
21.39	4.15	6
21.87	4.06	8
23.19	3.83	13
23.94	3.71	14
24.78	3.59	6
25.65	3.47	18
25.97	3.43	6
26.94	3.31	5
27.59	3.23	3
28.06	3.18	5
29.53	3.02	6

(B-iii) Compound (1) hydrochloride - Form FH3

Form FH3 was observed in precipitation experiments with ethanol or isopropanol solutions of form FH1 as well as in the degradation of form FH2. Form FH3 is stable in air and at 40°C and 75 % RH for at least one month. The preparation of form FH3 is described in Example 3 above. The XRPD pattern for form FH3 is shown in Figure 10 and the main peaks are listed in Table 10.

Table 10. Main XRPD peaks for Compound (1) Hydrochloride Form FH3

2 θ /°	d/Å	I/%
5.83	15.15	5
9.35	9.45	100
10.40	8.50	89
10.78	8.20	19
11.35	7.79	11
11.71	7.55	16
12.51	7.07	48
13.35	6.63	10
13.81	6.41	17
14.10	6.27	5
14.78	5.99	42

17.18	5.16	8
17.65	5.02	9
18.74	4.73	51
19.09	4.65	35
19.46	4.56	13
20.11	4.41	8
21.18	4.19	18
21.68	4.10	28
22.32	3.98	76
23.07	3.85	28
23.71	3.75	16
24.86	3.58	96
25.14	3.54	45
26.49	3.36	5
27.03	3.30	8
28.09	3.17	14
28.70	3.11	16
29.02	3.07	29
29.52	3.02	17

(B-iv) Compound (1) hydrochloride - Form FH4

Form FH4 was observed only in one precipitation experiment (DMF/dioxane). This form is unstable and disintegrates in air. A saturated solution of form FH1 in DMF was prepared at room temperature. Slow precipitation with approximately 4 volumes of 1,4-dioxane gave form FH4. The XRPD pattern of a fresh sample of FH4 is shown in Figure 11 and the main peaks are listed in Table 11 below. The sample disintegrated in air.

Table 11. Main XRPD peaks for Compound (1) hydrochloride - Form FH4

2 θ /°	d/Å	I/%
7.04	12.55	31
9.89	8.93	10
11.62	7.61	100
12.30	7.19	10
13.27	6.67	8
14.14	6.26	14
15.54	5.70	57
16.06	5.51	17
16.68	5.31	34
17.99	4.93	13
18.54	4.78	26
19.24	4.61	19
20.73	4.28	43
22.26	3.99	28
22.94	3.87	27
23.36	3.81	13
23.77	3.74	35

24.63	3.61	12
25.07	3.55	36
25.72	3.46	8
26.91	3.31	15
27.63	3.23	11

(B-v) Compound (1) hydrochloride - Form FH5

Form Fh5 was observed only in one precipitation experiment (methanol/acetone). This form is unstable and dissolves in moist air. A saturated solution of Form FH1 in methanol was prepared at room temperature. Slow precipitation with approximately 4 volumes of acetone gave form FH5. The XRPD pattern of a fresh sample of FH5 is shown in Figure 12 and the main peaks are listed in Table 12 below. The sample disintegrates in air.

Table 12. Main XRPD peaks for Compound (1) hydrochloride - Form FH5

2 θ /°	d/Å	I/%
2.32	38.00	100
6.15	14.35	18
11.79	7.50	6
15.79	5.61	5
20.81	4.27	8
22.76	3.90	3
23.76	3.74	5

10C. Compound (1) L-Lactate 1:1 salt crystal forms

(C-i) Compound (1) L-Lactate - Form FL1

10 The L-Lactate salt form FL1 was prepared as described in Example 2 above.

Form FL1 is stable in air and at 40°C and 75 % RH for at least one month. The XRPD pattern for form FL1 is shown in Figure 13 and the main peaks are listed in Table 13.

Table 13. Main XRPD peaks for Compound (1) Lactate – Form FL1

2 θ /°	d/Å	I/%
6.18	14.30	15
6.53	13.52	50
8.39	10.54	19
11.08	7.98	7
13.10	6.75	85
14.13	6.26	33
14.40	6.15	23
15.21	5.82	4
16.21	5.46	6
16.81	5.27	100
17.22	5.15	45
18.65	4.75	23

19.52	4.54	33
19.82	4.48	34
20.49	4.33	7
20.76	4.27	13
21.13	4.20	17
22.02	4.03	12
22.33	3.98	44
22.84	3.89	40
23.09	3.85	25
23.94	3.71	14
25.19	3.53	7
26.41	3.37	14
26.95	3.31	5
27.81	3.21	14

(C-ii) Compound (1) L-Lactate - Form FL2

Form FL2 was observed in precipitation experiments of methanol solutions of form FL1. Single crystal X-ray analysis showed that form FL2 is hydrated. It is nominally a tri-hydrate because there are 3 crystal water positions in the asymmetric unit, but they are not 100% occupied at room temperature and laboratory humidity. A saturated solution of form FL1 in methanol: water 9:1 was prepared at room temperature. Slow precipitation with approximately 4 volumes of acetone gave form FL2 which is stable in air. The XRPD pattern for form FL2 is shown in Figure 14 and the main peaks are listed in Table 14 below. A crystal packing diagram is shown in Figure 15 and the atom coordinates are listed in Table 15 below.

Table 14. Main XRPD peaks for Compound (1) Lactate salt – form FL2

2 θ /°	d/Å	I/%
8.03	11.00	29
10.71	8.26	53
11.98	7.38	90
13.13	6.74	49
15.39	5.75	29
16.09	5.50	32
16.61	5.33	42
17.26	5.13	37
18.17	4.88	20
18.82	4.71	56
20.40	4.35	40
21.01	4.22	49
21.53	4.12	27
22.34	3.98	100
22.56	3.94	73
23.71	3.75	82

24.30	3.66	8
24.65	3.61	12
26.56	3.35	13
27.70	3.22	21
28.29	3.15	16

Table 15. Unit cell parameters and coordinates in cif format for crystal structure of Compound (1) Lactate salt – form FL2

```

space group: P21
5  unit cell at 293K with a, b, c & β having 5% s.u.:
    a=5.8
    b=16.6
    c=14.9
    beta=98
10  alpha=gamma=90

Coordinates in cif format:

loop_
15  _atom_site_label
    _atom_site_type_symbol
    _atom_site_fract_x
    _atom_site_fract_y
    _atom_site_fract_z
20  _atom_site_U_iso_or_equiv
    _atom_site_adp_type
    _atom_site_occupancy
    _atom_site_symmetry_multiplicity
    _atom_site_calc_flag
25  _atom_site_refinement_flags
    _atom_site_disorder_assembly
    _atom_site_disorder_group
C1  C  -0.643(2)  1.1037(6)  0.6763(7)  0.097(3)  Uani  1  1  d  .  .  .
H1A H  -0.6995  1.0577  0.6395  0.117  Uiso  1  1  calc  .  .  .
30  H1B H  -0.5231  1.1308  0.6484  0.117  Uiso  1  1  calc  .  .  .
N2  N  -0.5563(16)  1.0791(5)  0.7694(6)  0.096(2)  Uani  1  1  d  .  .  .
C3  C  -0.692(3)  1.1148(8)  0.8352(8)  0.124(4)  Uani  1  1  d  .  .  .
H3A H  -0.7713  1.0734  0.8651  0.148  Uiso  1  1  calc  .  .  .
H3B H  -0.5925  1.1454  0.8805  0.148  Uiso  1  1  calc  .  .  .
35  C4  C  -0.8553(19)  1.1667(7)  0.7825(7)  0.094(3)  Uani  1  1  d  .  .  .
C5  C  -0.8393(19)  1.1609(6)  0.6900(7)  0.092(3)  Uani  1  1  d  .  .  .
C6  C  -1.036(3)  1.2141(8)  0.8083(8)  0.110(3)  Uani  1  1  d  .  .  .
H6  H  -1.0636  1.2139  0.8682  0.132  Uiso  1  1  calc  .  .  .
C7  C  -1.172(2)  1.2611(8)  0.7456(8)  0.105(3)  Uani  1  1  d  .  .  .
40  C8  C  -1.145(2)  1.2560(8)  0.6564(9)  0.111(3)  Uani  1  1  d  .  .  .
H8  H  -1.2387  1.2867  0.6138  0.133  Uiso  1  1  calc  .  .  .
C9  C  -0.979(2)  1.2053(9)  0.6287(7)  0.109(3)  Uani  1  1  d  .  .  .
H9  H  -0.9640  1.2017  0.5677  0.130  Uiso  1  1  calc  .  .  .
C10 C  -1.3561(18)  1.3173(8)  0.7739(9)  0.106(3)  Uani  1  1  d  .  .  .
45  H10A H  -1.4455  1.3402  0.7202  0.127  Uiso  1  1  calc  .  .  .
    H10B H  -1.4617  1.2864  0.8055  0.127  Uiso  1  1  calc  .  .  .

```

	N11	N	-1.2550(14)	1.3836(6)	0.8332(6)	0.096(2)	Uani	1	1	d	.	.	.
	C12	C	-1.1136(17)	1.4353(6)	0.7839(7)	0.091(3)	Uani	1	1	d	.	.	.
	H12A	H	-1.2098	1.4591	0.7324	0.109	Uiso	1	1	calc	.	.	.
	H12B	H	-0.9935	1.4035	0.7615	0.109	Uiso	1	1	calc	.	.	.
5	C13	C	-1.0015(17)	1.5021(7)	0.8462(8)	0.100(3)	Uani	1	1	d	.	.	.
	H13A	H	-0.8991	1.4783	0.8961	0.121	Uiso	1	1	calc	.	.	.
	H13B	H	-0.9092	1.5368	0.8128	0.121	Uiso	1	1	calc	.	.	.
	N14	N	-1.1853(15)	1.5509(5)	0.8822(6)	0.094(2)	Uani	1	1	d	.	.	.
	H14	H	-1.2741	1.5755	0.8352	0.113	Uiso	1	1	calc	.	.	.
10	C15	C	-1.3350(18)	1.4966(7)	0.9279(7)	0.095(3)	Uani	1	1	d	.	.	.
	H15A	H	-1.4599	1.5276	0.9479	0.114	Uiso	1	1	calc	.	.	.
	H15B	H	-1.2441	1.4730	0.9808	0.114	Uiso	1	1	calc	.	.	.
	C16	C	-1.4358(17)	1.4308(7)	0.8658(8)	0.098(3)	Uani	1	1	d	.	.	.
	H16A	H	-1.5310	1.3959	0.8977	0.117	Uiso	1	1	calc	.	.	.
15	H16B	H	-1.5346	1.4542	0.8148	0.117	Uiso	1	1	calc	.	.	.
	C17	C	-1.068(2)	1.6140(9)	0.9439(9)	0.119(4)	Uani	1	1	d	.	.	.
	H17A	H	-1.1835	1.6447	0.9694	0.178	Uiso	1	1	calc	.	.	.
	H17B	H	-0.9807	1.6492	0.9103	0.178	Uiso	1	1	calc	.	.	.
	H17C	H	-0.9658	1.5886	0.9916	0.178	Uiso	1	1	calc	.	.	.
20	C18	C	-0.382(2)	1.0287(9)	0.7999(8)	0.113(4)	Uani	1	1	d	.	.	.
	O19	O	-0.345(2)	1.0216(8)	0.8837(6)	0.156(4)	Uani	1	1	d	.	.	.
	C20	C	-0.228(2)	0.9847(6)	0.7418(7)	0.096(3)	Uani	1	1	d	.	.	.
	C21	C	-0.069(3)	0.9286(9)	0.7863(9)	0.119(4)	Uani	1	1	d	.	.	.
	C22	C	0.064(2)	0.8867(9)	0.7367(9)	0.114(4)	Uani	1	1	d	.	.	.
25	H22	H	0.1812	0.8547	0.7669	0.137	Uiso	1	1	calc	.	.	.
	C23	C	0.038(2)	0.8879(7)	0.6447(8)	0.097(3)	Uani	1	1	d	.	.	.
	C24	C	-0.1201(18)	0.9425(7)	0.5972(8)	0.096(3)	Uani	1	1	d	.	B	.
	C25	C	-0.253(2)	0.9882(7)	0.6463(8)	0.100(3)	Uani	1	1	d	.	.	.
	H25	H	-0.3632	1.0228	0.6160	0.120	Uiso	1	1	calc	.	.	.
30	O26	O	-0.036(2)	0.9229(9)	0.8775(6)	0.169(5)	Uani	1	1	d	.	.	.
	H26	H	-0.1427	0.9456	0.8980	0.253	Uiso	1	1	calc	R	.	.
	O27	O	0.1658(15)	0.8404(5)	0.5948(6)	0.118(3)	Uani	1	1	d	.	.	.
	H27	H	0.2091	0.7999	0.6238	0.176	Uiso	1	1	calc	R	.	.
	C28	C	-0.141(4)	0.9478(11)	0.4948(10)	0.138(6)	Uani	1	1	d	.	.	.
35	H28	H	-0.0894	0.8953	0.4750	0.166	Uiso	1	1	calc	.	A	1
	C29	C	-0.029(11)	1.004(4)	0.449(3)	0.24(3)	Uani	0.58(6)	1	d	P	B	1
	H29A	H	-0.0741	0.9976	0.3847	0.363	Uiso	0.58	1	calc	P	B	1
	H29B	H	0.1361	0.9972	0.4628	0.363	Uiso	0.58	1	calc	P	B	1
	H29C	H	-0.0703	1.0575	0.4662	0.363	Uiso	0.58	1	calc	P	B	1
40	C30	C	-0.417(7)	0.950(3)	0.4621(19)	0.159(19)	Uani	0.58(6)	1	d	P	B	1
	H30A	H	-0.4911	0.9083	0.4918	0.239	Uiso	0.58	1	calc	P	B	1
	H30B	H	-0.4462	0.9424	0.3978	0.239	Uiso	0.58	1	calc	P	B	1
	H30C	H	-0.4773	1.0016	0.4772	0.239	Uiso	0.58	1	calc	P	B	1
45	C29	C	-0.156(11)	1.040(2)	0.465(2)	0.14(2)	Uani	0.42(6)	1	d	P	B	2
	H29D	H	-0.0071	1.0655	0.4814	0.215	Uiso	0.42	1	calc	P	B	2
	H29E	H	-0.2703	1.0675	0.4943	0.215	Uiso	0.42	1	calc	P	B	2
	H29F	H	-0.1983	1.0438	0.4003	0.215	Uiso	0.42	1	calc	P	B	2
	C30	C	-0.295(12)	0.897(4)	0.446(2)	0.150(19)	Uani	0.42(6)	1	d	P	B	2
50	H30D	H	-0.3403	0.9185	0.3870	0.224	Uiso	0.42	1	calc	P	B	2
	H30E	H	-0.4300	0.8910	0.4766	0.224	Uiso	0.42	1	calc	P	B	2
	H30F	H	-0.2234	0.8451	0.4418	0.224	Uiso	0.42	1	calc	P	B	2
	O1L	O	-1.5549(12)	1.6174(6)	0.7786(6)	0.124(3)	Uani	1	1	d	.	.	.
	O2L	O	-1.7419(12)	1.7087(6)	0.6890(7)	0.125(3)	Uani	1	1	d	.	.	.

C1L C -1.5569(17) 1.6742(7) 0.7238(8) 0.098(3) Uani 1 1 d . . .
 C2L C -1.3365(17) 1.6989(8) 0.6926(9) 0.108(4) Uani 1 1 d . . .
 H2L H -1.3065 1.7549 0.7117 0.129 Uiso 1 1 calc . . .
 C3L C -1.355(2) 1.6971(12) 0.5917(11) 0.143(5) Uani 1 1 d . . .
 5 H3L1 H -1.2130 1.7162 0.5734 0.214 Uiso 1 1 calc . . .
 H3L2 H -1.4813 1.7312 0.5662 0.214 Uiso 1 1 calc . . .
 H3L3 H -1.3842 1.6429 0.5706 0.214 Uiso 1 1 calc . . .
 O3L O -1.1538(13) 1.6538(7) 0.7316(8) 0.150(4) Uani 1 1 d . . .
 H3L H -1.0243 1.6711 0.7191 0.224 Uiso 1 1 d . . .
 10 O1W O -0.448(6) 1.237(6) 1.045(2) 0.45(5) Uani 0.78(6) 1 d P . .
 O2W O 0.021(15) 0.8037(17) 0.9990(19) 0.74(7) Uani 1 1 d . . .
 O3W O -0.35(3) 0.773(9) 0.953(15) 0.77(8) Uani 0.22(6) 1 d P . .

(C-iii) Compound (1) L-Lactate - Form FL3

15 Form FL3 was observed in precipitation experiments of THF solutions of form FL1. Form
 FL3 transforms in air into form FL1. A saturated solution of form FL1 in THF was prepared
 at room temperature. Slow precipitation with approximately 4 volumes of heptane gave
 form FL3. The XRPD pattern of a fresh sample of form FL3 is shown in Figure 16 and the
 main peaks are listed in Table 16 below. A sample of FL3 was dried in air for 2 days after
 20 which XRPD analysis showed that conversion to form FL1 had occurred.

Table 16. Main XRPD peaks for Compound (1) Lactate salt – form FL3

2 θ /°	d/Å	I/%
5.53	15.98	100
8.36	10.56	5
11.07	7.98	41
13.16	6.72	12
13.85	6.39	8
16.69	5.31	39
17.17	5.16	21
18.00	4.92	49
18.49	4.80	11
19.28	4.60	14
19.79	4.48	5
20.34	4.36	7
21.05	4.22	21
21.47	4.14	7
21.93	4.05	4
22.47	3.95	16
22.84	3.89	23
24.56	3.62	4
26.28	3.39	6
27.06	3.29	3
27.47	3.24	3

29.11	3.07	6
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10D. Compound (1) Sulphate 1:1 salt crystal forms

(D-i) Compound (1) Sulphate - Form FS1

Form FS1 was observed in crystallization experiments involving acetonitrile as the precipitant. It is unstable on air and transforms to form FS3. A saturated solution of the 1:1 salt of Compound (1) (prepared by dissolving Compound (1) in H₂SO₄ and evaporating to dryness) in water was prepared at room temperature. Slow precipitation with approximately 4 volumes of acetonitrile gave form FS1. The XRPD pattern for FS1 is shown in Figure 17 and the main peaks are listed in Table 17.

Table 17. Main XRPD peaks for Compound (1) Sulphate – Form FS1

2θ/°	d/Å	I/%
4.79	18.45	100
10.02	8.82	28
10.68	8.28	3
11.28	7.84	10
12.89	6.86	6
14.38	6.15	34
15.27	5.80	12
16.91	5.24	17
17.64	5.02	7
18.29	4.85	11
18.86	4.70	3
19.28	4.60	4
20.12	4.41	10
20.82	4.26	8
21.21	4.19	3
21.76	4.08	10
22.32	3.98	13
22.89	3.88	7
23.83	3.73	5
24.22	3.67	3
24.42	3.64	3
25.13	3.54	8
29.04	3.07	8

Compound (1) Sulphate - Form FS2

Form FS2 is unstable on air and transforms to form FS5. If kept at 40 °C and 75% RH, form FS2 transforms to form FS4. Compound (1) was dissolved in 1 mol equivalent of concentrated H₂SO₄, precipitated with approximately 4 volumes of acetonitrile and the

crystalline mass which formed was filtered. The XRPD pattern for FS2 is shown on Figure 18 and the main peaks are listed in Table 18.

Table 18. Main XRPD peaks for Compound (1) Sulphate – Form FS2

2 θ /°	d/Å	I/%
4.17	21.20	2
7.03	12.57	24
7.43	11.89	100
8.09	10.92	11
8.67	10.19	90
9.27	9.54	17
9.65	9.16	19
10.41	8.49	7
10.98	8.05	6
11.76	7.52	31
12.53	7.06	5
13.84	6.40	26
14.55	6.08	8
15.39	5.75	16
16.24	5.45	4
16.89	5.25	7
17.50	5.06	26
18.05	4.91	17
18.93	4.68	16
19.47	4.56	16
23.20	3.83	24
24.21	3.67	19
25.21	3.53	10
25.75	3.46	14
26.62	3.35	13
27.67	3.22	13

5 Compound (1) Sulphate - Form FS3

Form FS3 is a stable form that was observed as product of form FS1 after ageing in air and the transformation of form FS6 in a warm and humid environment (40°C, 75% RH). The XRPD pattern of a sample of form FS3 that was prepared by allowing form FS1 to dry for 2 days in air is shown in Figure 19 and the main peaks are listed in Table 19.

10 Table 19. Main XRPD peaks for Compound (1) Sulphate – Form FS3

2 θ /°	d/Å	I/%
4.81	18.36	17
5.43	16.25	100
10.30	8.58	48
11.24	7.87	24
12.94	6.84	5
13.98	6.33	7
14.26	6.21	26

14.91	5.94	33
15.62	5.67	12
16.41	5.40	56
17.53	5.05	26
18.38	4.82	28
18.61	4.76	40
19.01	4.66	22
19.38	4.58	10
19.92	4.45	30
20.27	4.38	13
20.71	4.28	6
21.19	4.19	9
21.77	4.08	31
22.67	3.92	20
23.79	3.74	19
24.23	3.67	27
25.36	3.51	21
27.38	3.25	6
28.82	3.09	9

Compound (1) Sulphate - Form FS4

Form FS4 is a stable form that was observed only as a product of the transformation of form FS2 in a warm and humid environment (40°C, 75%RH). An XRPD pattern of a sample of form FS4 that was prepared by incubating form FS2 for several weeks at 40°C and 75% RH is shown in Figure 20 and the main peaks are listed in Table 20.

Table 20. Main XRPD peaks for Compound (1) Sulphate – Form FS4

2 θ /°	d/Å	I/%
4.64	19.03	4
7.16	12.34	39
7.48	11.80	100
7.97	11.08	29
8.42	10.49	13
8.82	10.02	34
9.09	9.73	29
9.37	9.43	35
10.45	8.46	30
11.77	7.51	53
13.25	6.68	17
13.54	6.54	16
14.36	6.16	24
15.03	5.89	13
16.21	5.46	21
16.99	5.22	33
17.28	5.13	31
17.59	5.04	30
17.96	4.93	19
18.90	4.69	24

19.43	4.57	10
19.83	4.47	8
21.36	4.16	12
23.13	3.84	31
23.68	3.75	28
23.96	3.71	32
24.77	3.59	18
25.64	3.47	17
26.19	3.40	14
26.73	3.33	13
27.20	3.28	11
27.76	3.21	17
28.64	3.11	9

Compound (1) Sulphate - Form FS5

Form FS53 is a stable form that was observed as a product of the ageing of form FS2 in air and the transformation of form FS4 in a dry environment (20°C, 11% RH). The XRPD pattern of a sample of form FS5 that was prepared by allowing form FS2 to dry for 2 days in air is shown in Figure 21 and the main peaks are listed in Table 21.

Table 21. Main XRPD peaks for Compound (1) Sulphate – Form FS5

2 θ /°	d/Å	I/%
4.70	18.80	19
7.11	12.42	56
7.99	11.05	100
9.33	9.47	42
9.57	9.23	29
10.45	8.46	54
11.64	7.60	30
13.27	6.67	62
14.28	6.20	20
14.65	6.04	13
15.12	5.86	14
15.60	5.67	24
16.98	5.22	88
17.65	5.02	22
18.01	4.92	35
18.80	4.72	48
19.32	4.59	17
19.83	4.47	13
21.08	4.21	18
23.21	3.83	39
23.51	3.78	23
23.92	3.72	36
24.30	3.66	19
25.06	3.55	22
26.24	3.39	37
27.28	3.27	13

28.67	3.11	18
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Compound (1) Sulphate - Form FS6

Form FS6 was identified in a number of different crystallization experiments during form screening. It is stable in air but, in a warm and humid environment (40°C, 75% RH), it
5 transforms into form FS3.

A saturated solution of the 1:1 sulphate salt of Compound (1) (prepared by dissolving Compound (1) in H₂SO₄ and evaporating to dryness) in DMF was prepared at room temperature. Slow precipitation with approximately 4 volumes of toluene gave form FS6. The XRPD pattern for FS6 is shown in Figure 22 and the main peaks are listed in Table 22.

10 Table 22. Main XRPD peaks for Compound (1) Sulphate – Form FS6

2θ/°	d/Å	I/%
4.82	18.32	100
9.98	8.86	32
11.35	7.79	9
12.92	6.85	4
14.45	6.13	36
15.38	5.76	17
16.97	5.22	19
17.52	5.06	7
18.18	4.87	15
19.42	4.57	9
20.23	4.39	16
20.93	4.24	13
21.31	4.17	5
21.66	4.10	5
21.89	4.06	7
22.29	3.98	17
22.84	3.89	8
23.04	3.86	6
23.94	3.71	4
24.51	3.63	4
25.26	3.52	9
29.18	3.06	7

EXAMPLE 11

Determination of Antifungal Activity

The antifungal activity of the compounds of the formula (I) is determined using the following
15 protocol.

The compounds are tested against a panel of fungi including *Candida parapsilosis*, *Candida tropicalis*, *Candida albicans*-ATCC 36082 and *Cryptococcus neoformans*. The

test organisms are maintained on Sabourahd Dextrose Agar slants at 4 °C. Singlet suspensions of each organism are prepared by growing the yeast overnight at 27 °C on a rotating drum in yeast-nitrogen base broth (YNB) with amino acids (Difco, Detroit, Mich.), pH 7.0 with 0.05 morpholine propanesulphonic acid (MOPS). The suspension is then

5 centrifuged and washed twice with 0.85% NaCl before sonicating the washed cell suspension for 4 seconds (Branson Sonifier, model 350, Danbury, Conn.). The singlet blastospores are counted in a haemocytometer and adjusted to the desired concentration in 0.85% NaCl.

The activity of the test compounds is determined using a modification of a broth

10 microdilution technique. Test compounds are diluted in DMSO to a 1.0 mg/ml ratio then diluted to 64 µg/ml in YNB broth, pH 7.0 with MOPS (Fluconazole is used as the control) to provide a working solution of each compound. Using a 96-well plate, wells 1 and 3 through 12 are prepared with YNB broth, ten fold dilutions of the compound solution are made in wells 2 to 11 (concentration ranges are 64 to 0.125 µg/ml). Well 1 serves as a sterility

15 control and blank for the spectrophotometric assays. Well 12 serves as a growth control. The microtitre plates are inoculated with 10 µl in each of well 2 to 11 (final inoculum size is 10⁴ organisms/ml). Inoculated plates are incubated for 48 hours at 35 °C. The MIC values are determined spectrophotometrically by measuring the absorbance at 420 nm (Automatic Microplate Reader, DuPont Instruments, Wilmington, Del.) after agitation of the plates for 2

20 minutes with a vortex-mixer (Vorte-Genie 2 Mixer, Scientific Industries, Inc., Bolemia, N.Y.). The MIC endpoint is defined as the lowest drug concentration exhibiting approximately 50% (or more) reduction of the growth compared with the control well. With the turbidity assay this is defined as the lowest drug concentration at which turbidity in the well is <50% of the control (IC₅₀). Minimal Cytolytic Concentrations (MCC) are determined by sub-

25 culturing all wells from the 96-well plate onto a Sabourahd Dextrose Agar (SDA) plate, incubating for 1 to 2 days at 35 °C and then checking viability.

EXAMPLE 12

Methods Of Testing For Pain Reducing Or Pain Preventing Activity

(I) Inflammatory hyperalgesia test

30 Mechanical hyperalgesia can be examined in a rat model of inflammatory pain. Paw withdrawal thresholds to an increasing pressure stimulus are measured by the Randal-Sellito technique using an analgesymeter (Ugo Basile, Milan), in naïve animals prior to an intraplantar injection of complete Freund's complete adjuvant (FCA) into the left hind paw. 24 hours later paw withdrawal thresholds are measured again prior to (predose) and then

from 10 minutes to 6 hours following drug or vehicle administration. Reversal of hyperalgesia in the ipsilateral paw is calculated according to the formula:

$$\% \text{ reversal} = \frac{\text{postdose threshold} - \text{predose threshold}}{\text{naive threshold} - \text{predose threshold}} \times 100$$

(ii) Neuropathic hyperalgesia test

- 5 Mechanical hyperalgesia can be examined in a rat model of neuropathic pain induced by partial ligation of the left sciatic nerve. Approximately 14 days following surgery mechanical withdrawal thresholds of both the ligated (ipsilateral) and non-ligated (contralateral) paw are measured prior to (predose) and then from 10 minutes to 6 hours following drug or vehicle administration. Reversal of hyperalgesia at each time point is
10 calculated according to the formula:

$$\% \text{ reversal} = \frac{\text{ipsilateral threshold postdose} - \text{ipsilateral threshold predose}}{\text{contralateral threshold predose} - \text{ipsilateral threshold predose}} \times 100$$

- All experiments are carried out using groups of 6 animals. Stock concentrations of drugs are dissolved in distilled water and subsequent dilutions were made in 0.9% saline for subcutaneous administration in a volume of 4 mlkg⁻¹. All drugs are made up in plastic vials
15 and kept in the dark.

Statistical analysis are carried out on withdrawal threshold readings (g) using ANOVA with repeated measures followed by Tukey's HSD test. Efficacy refers to the maximal reversal of hyperalgesia observed at the doses used.

(iii) Testing the effects of compounds of formula (0) a Rat Model of Bone Cancer Pain

- 20 Adult female rats are given intra-tibial injections of MRMZ-1 rat mammary gland carcinoma cells (3 µl, 10⁷ cells/ml). The animals typically gradually develop mechanical hyperalgesia, mechanical allodynia (skin sensitivity to non-noxious stimuli) and hind limb sparing, beginning on day 12-14 following cell injection. A compound of formula (0) (e.g. at a dose of 10 and 30 µg/kg s.c.) is administered 3 times a week from the day of cell injection, and
25 the extent of inhibition of hind limb sparing and mechanical allodynia is determined in comparison to vehicle-treated controls.

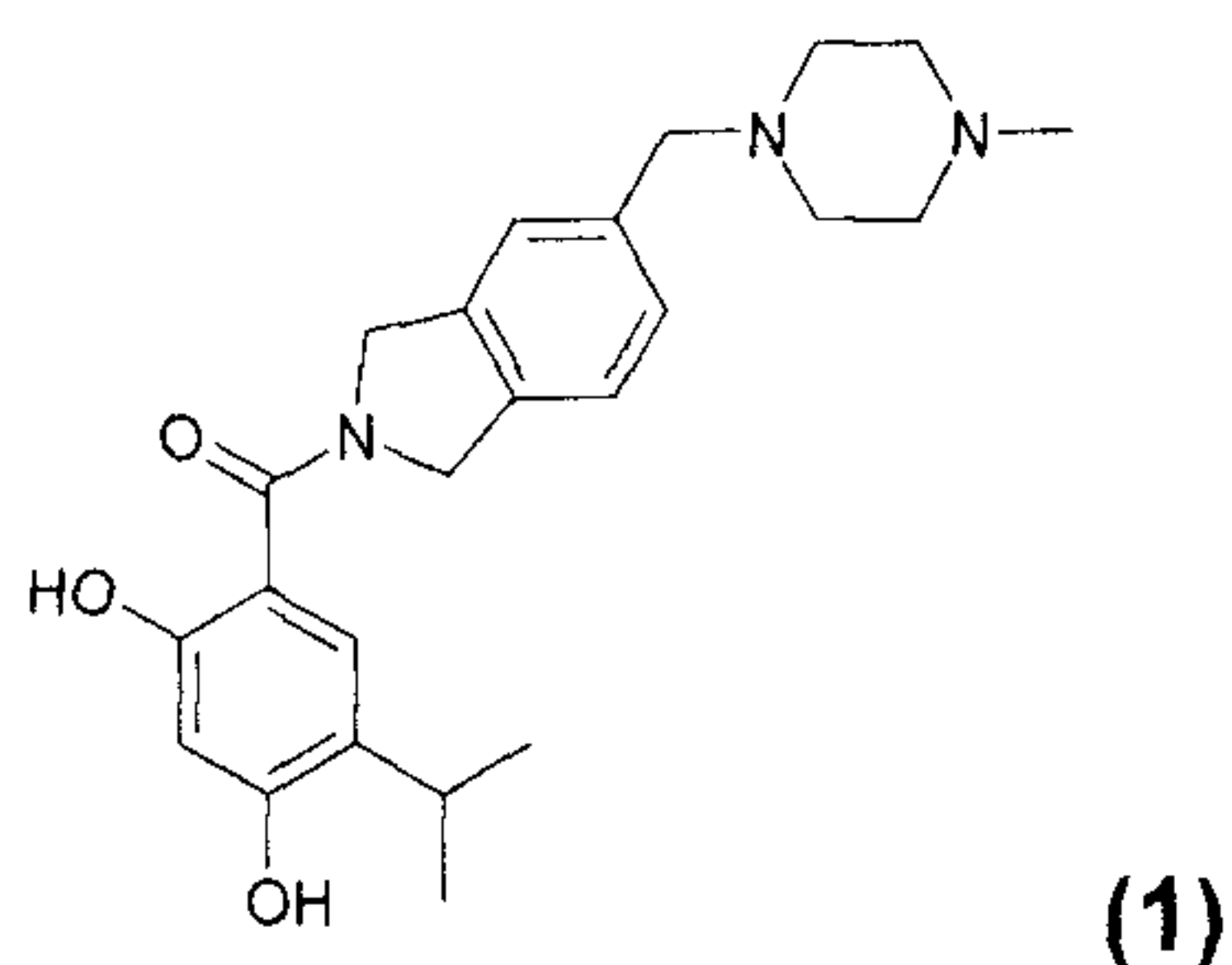
Equivalents

- The foregoing examples are presented for the purpose of illustrating the invention and should not be construed as imposing any limitation on the scope of the invention. It will
30 readily be apparent that numerous modifications and alterations may be made to the specific embodiments of the invention described above and illustrated in the examples

without departing from the principles underlying the invention. All such modifications and alterations are intended to be embraced by this application.

CLAIMS

1. An acid addition salt of a compound of the formula **(1)**



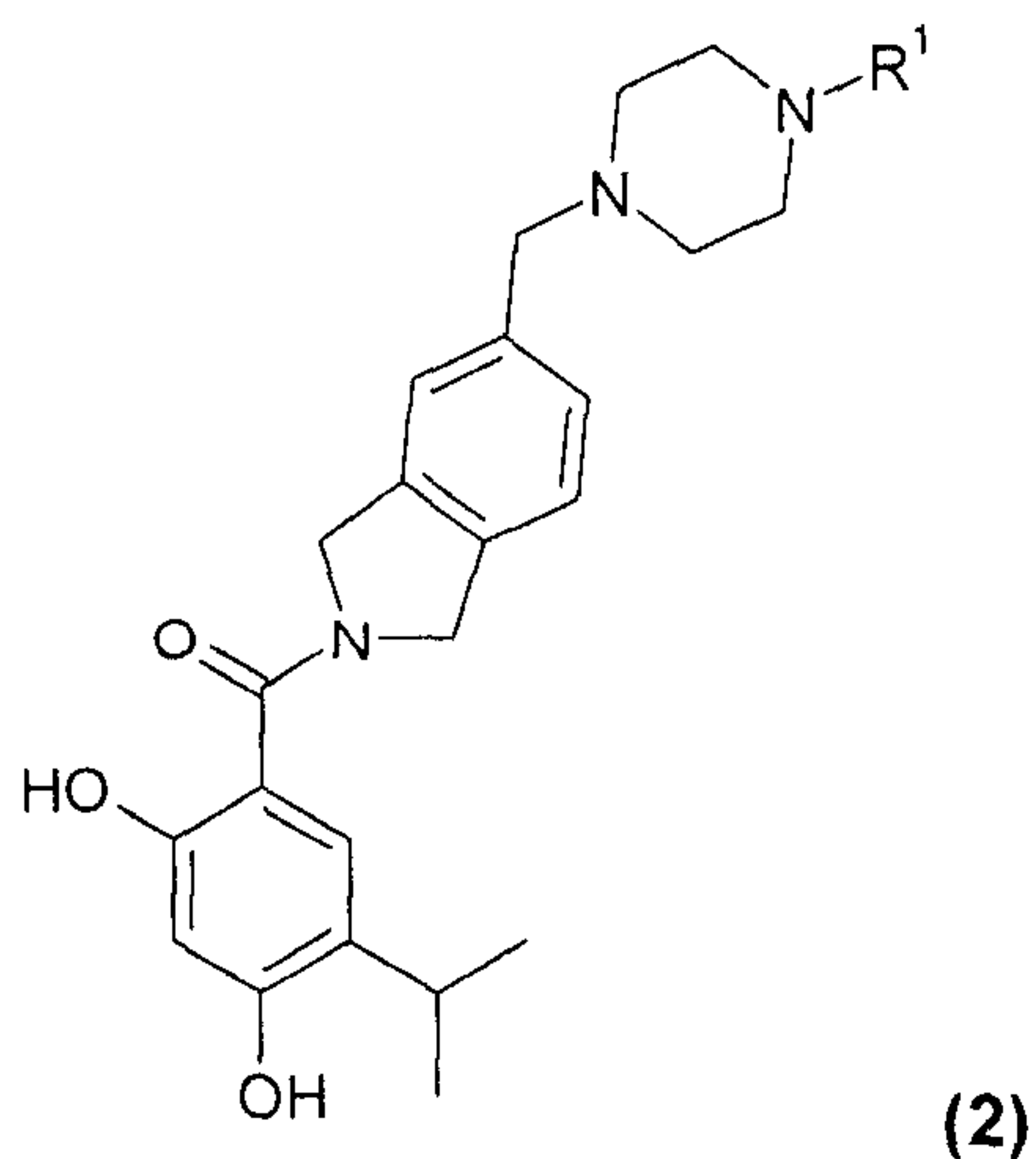
which is a salt formed with lactic acid.

2. An acid addition salt according to claim 1 which is a salt formed with L-lactic acid.
3. An acid addition salt in crystalline form according to claim 1 wherein the crystalline form is selected from forms:

FL1: characterised by an XRPD pattern having a diffraction angle ($2\theta/^\circ$) peak at 16.81, and

FL2: characterised by an XRPD pattern having a diffraction angle ($2\theta/^\circ$) peak at 22.34.

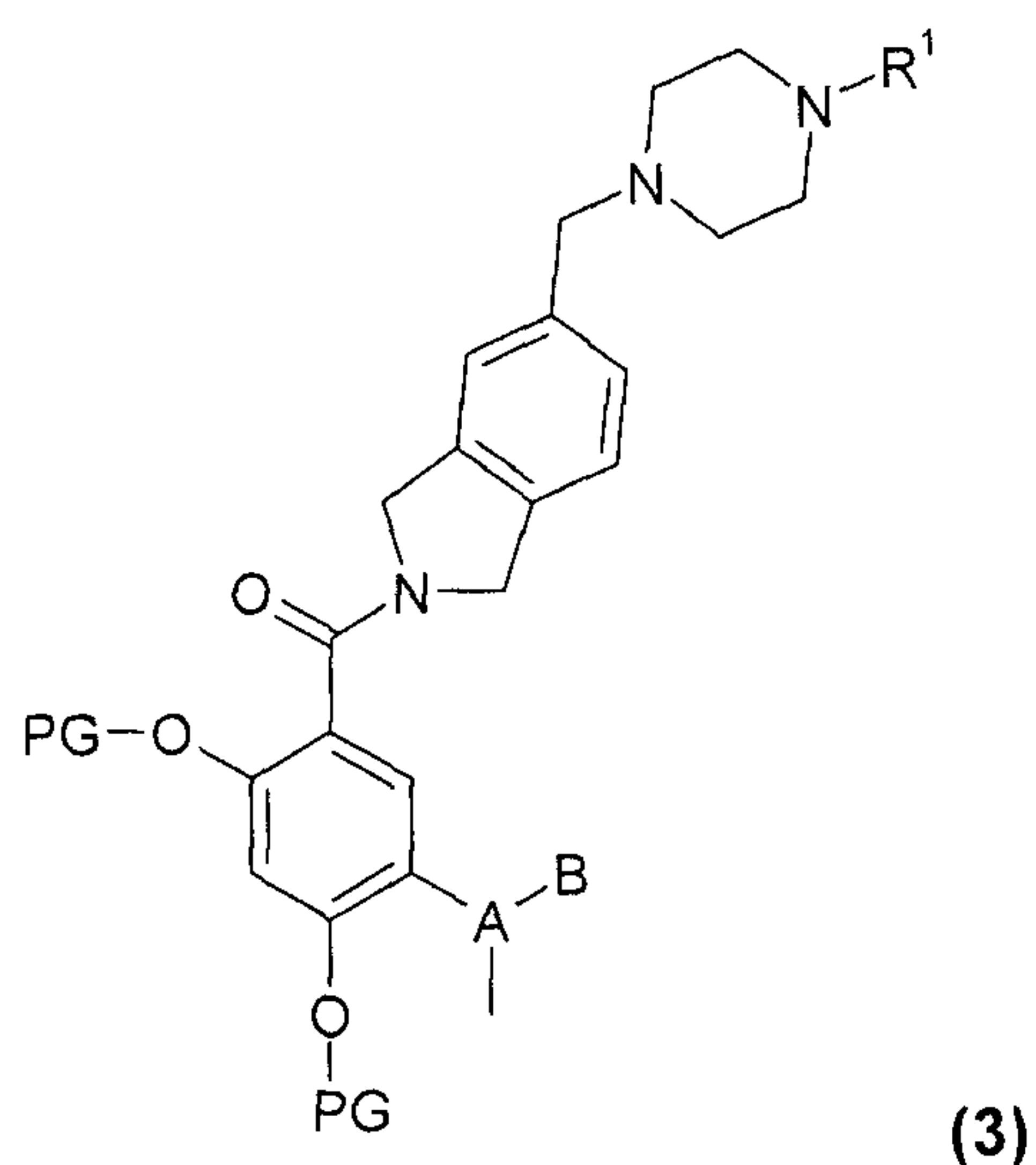
4. A process for the preparation of a compound of the formula **(2)**:



wherein R^1 is C_{1-4} alkyl; which process comprises:

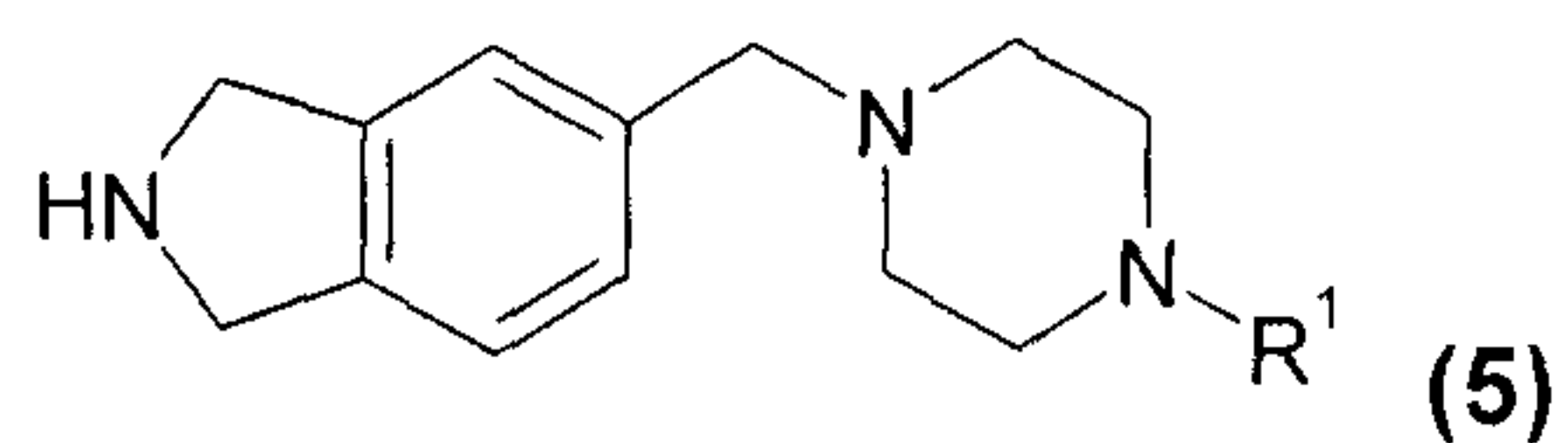
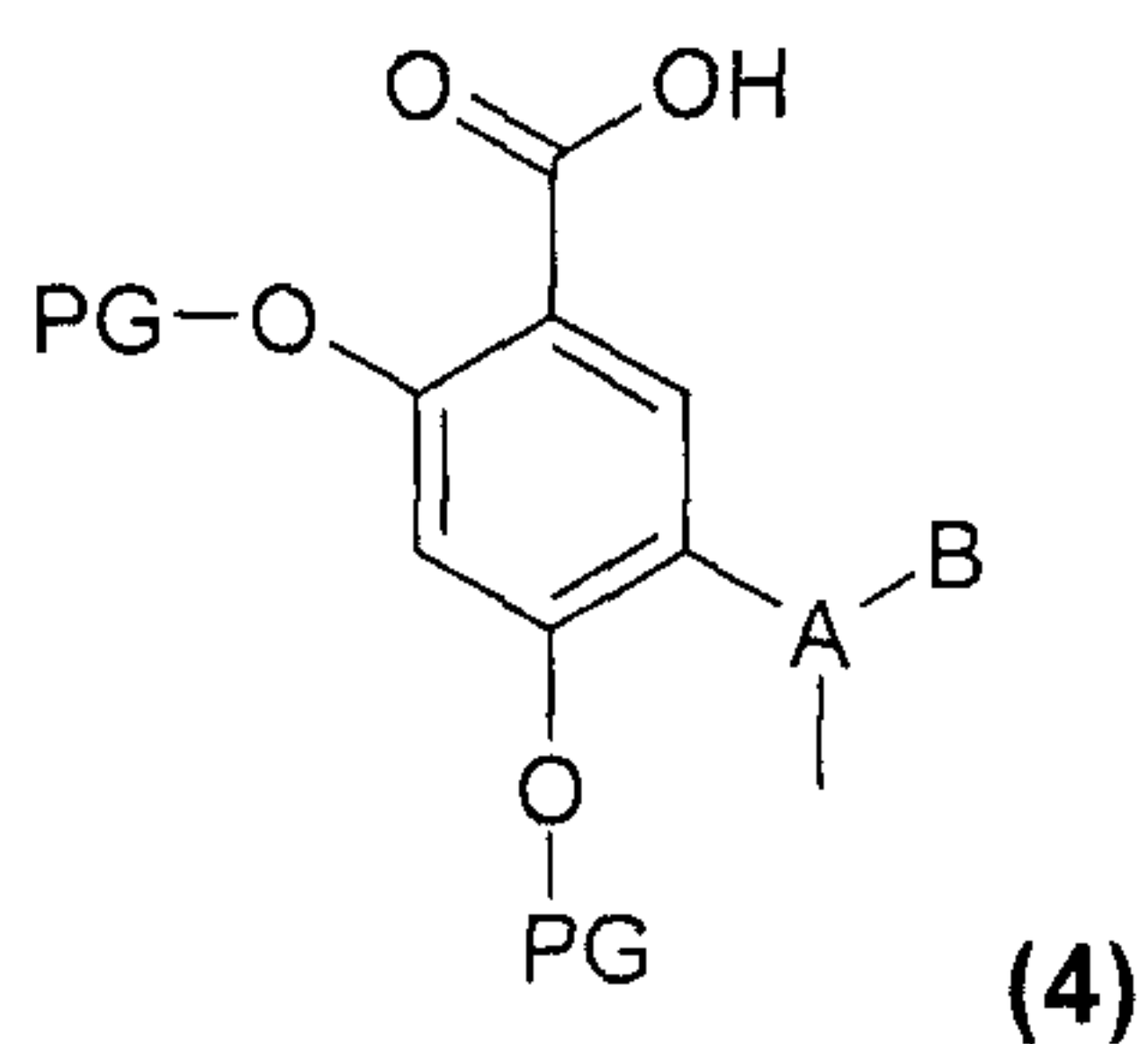
either subjecting to catalytic hydrogenation a compound of the formula **(3)**:

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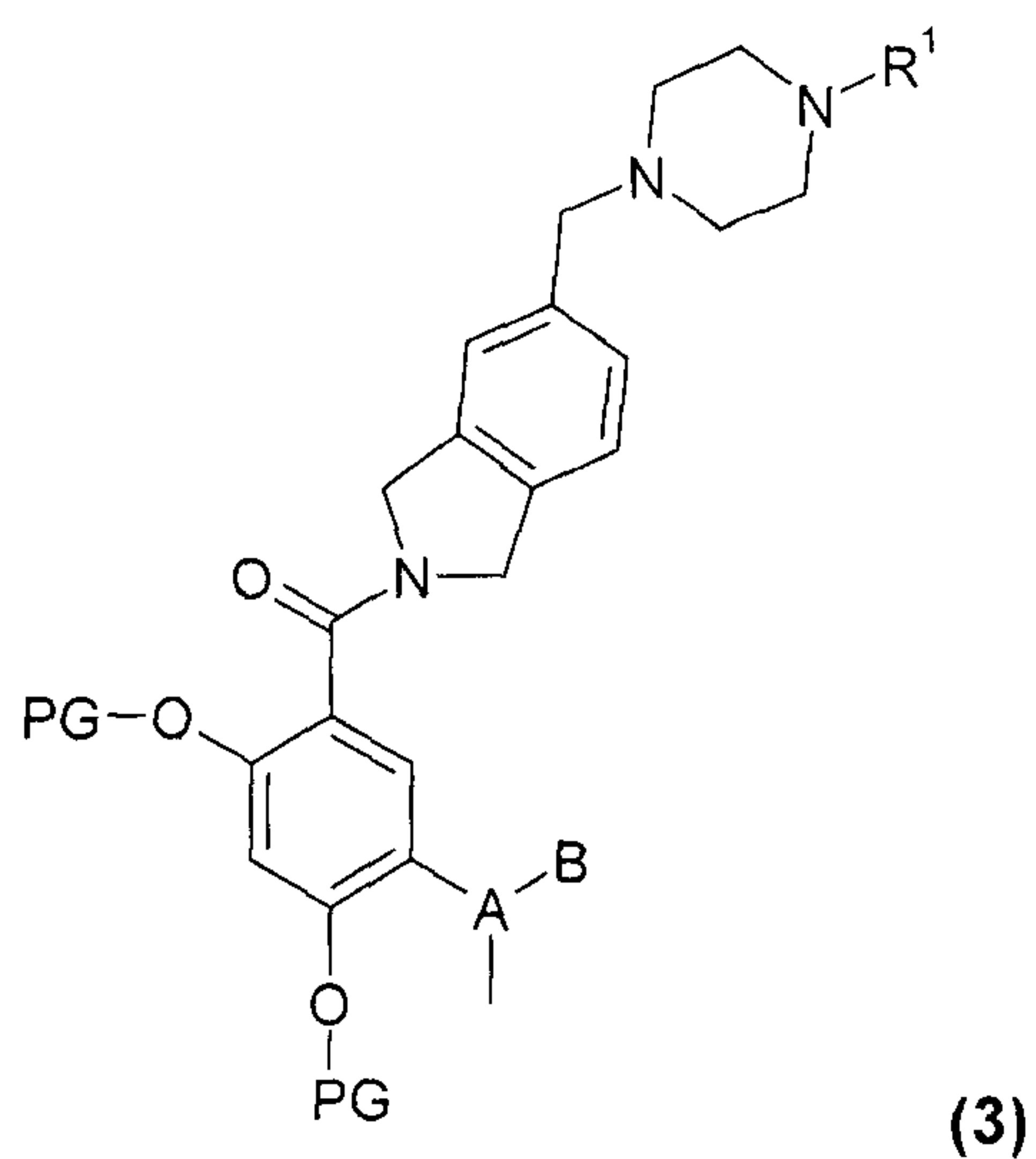


wherein PG is a protecting group removable under hydrogenation conditions and A-B is $C=CH_2$, and thereafter, where the compound of formula (2) is prepared in the form of a free base, optionally converting the free base to an acid addition salt; or

(a-i) the reaction of a compound of the formula (4), or an activated form or derivative thereof with a compound of the formula (5):



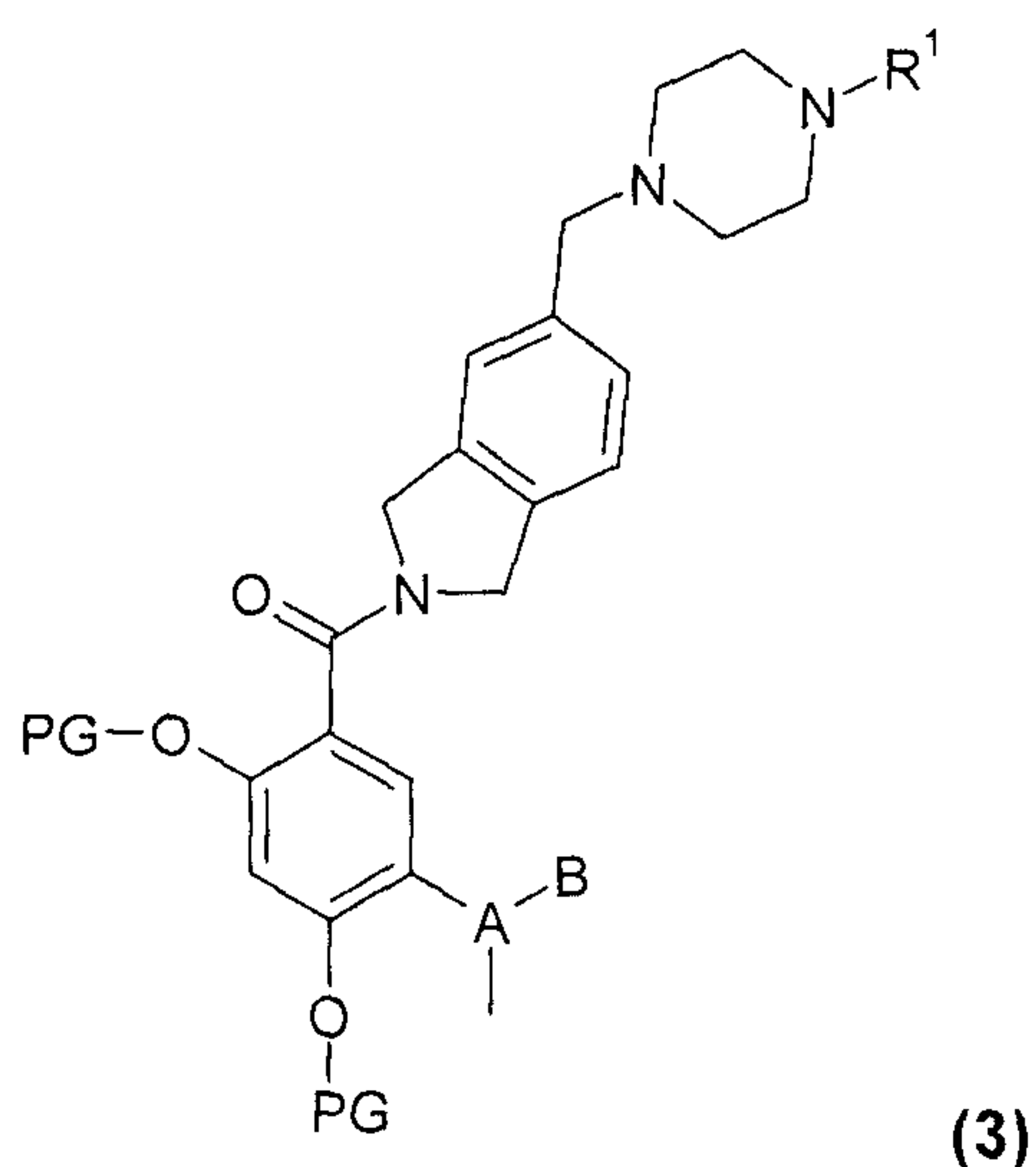
under amide forming conditions to give a compound of the formula (3);



wherein R^1 , PG, A-B are as defined above; and

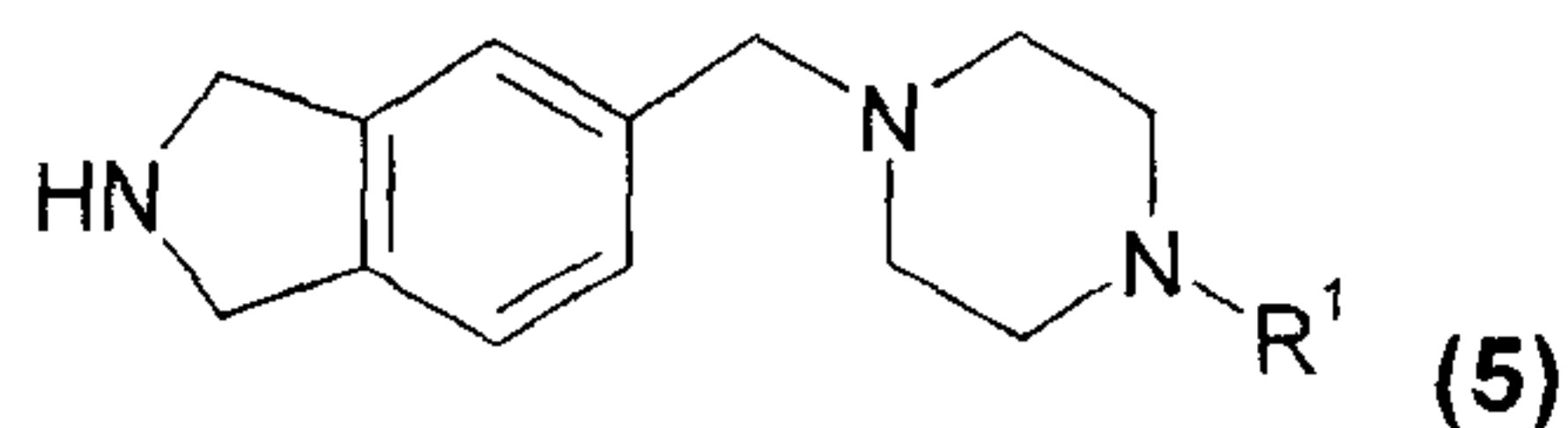
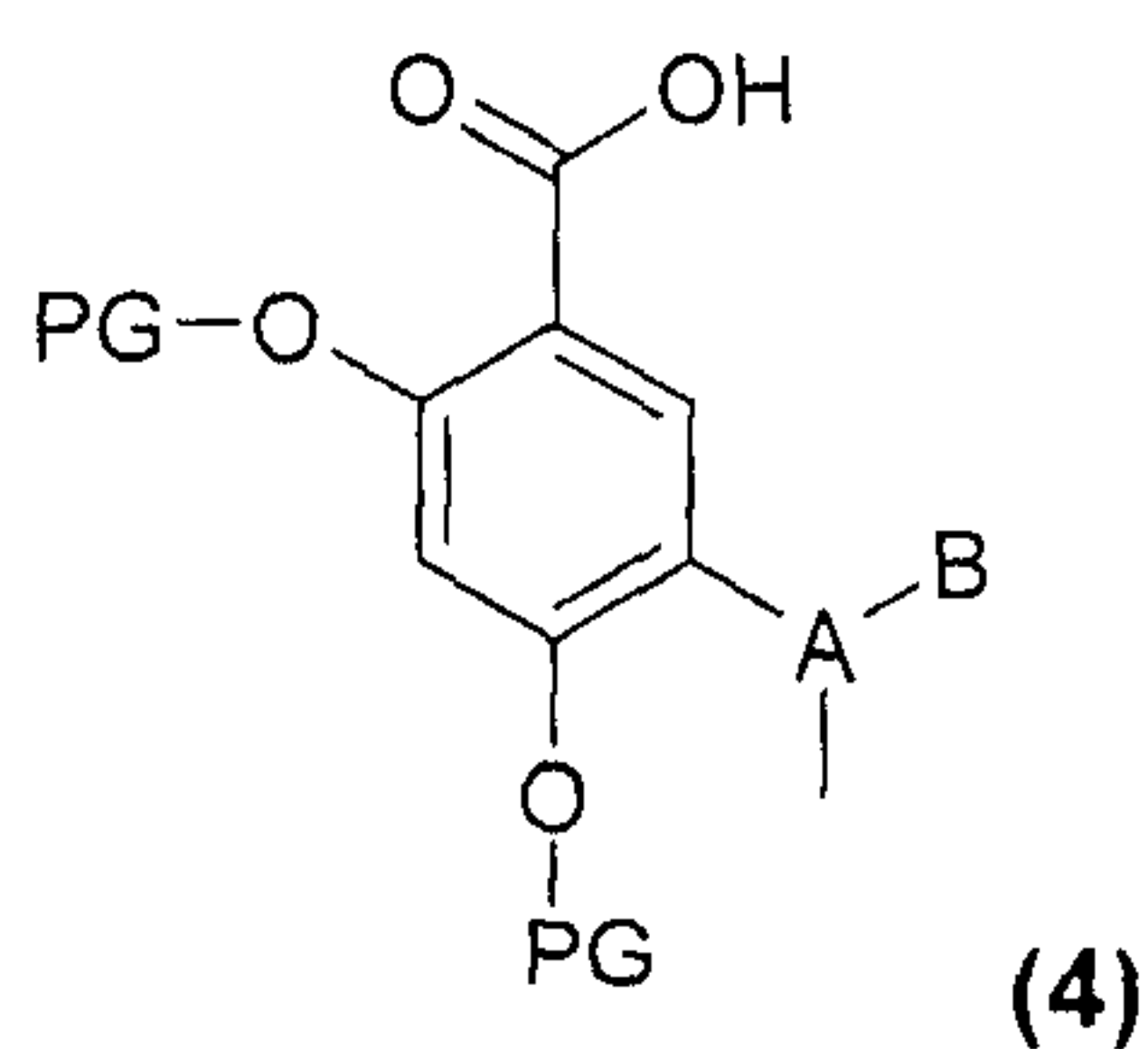
(b) subjecting the compound of formula (3) to catalytic hydrogenation to remove the protecting groups PG and to reduce the group A-B to an isopropyl group and thereafter, where the compound of formula (2) is prepared in the form of a free base, optionally converting the free base to an acid addition salt.

5. A process for the preparation of a compound of the formula (3):



wherein R¹, PG, A-B are as defined in claim 4; which process comprises either:

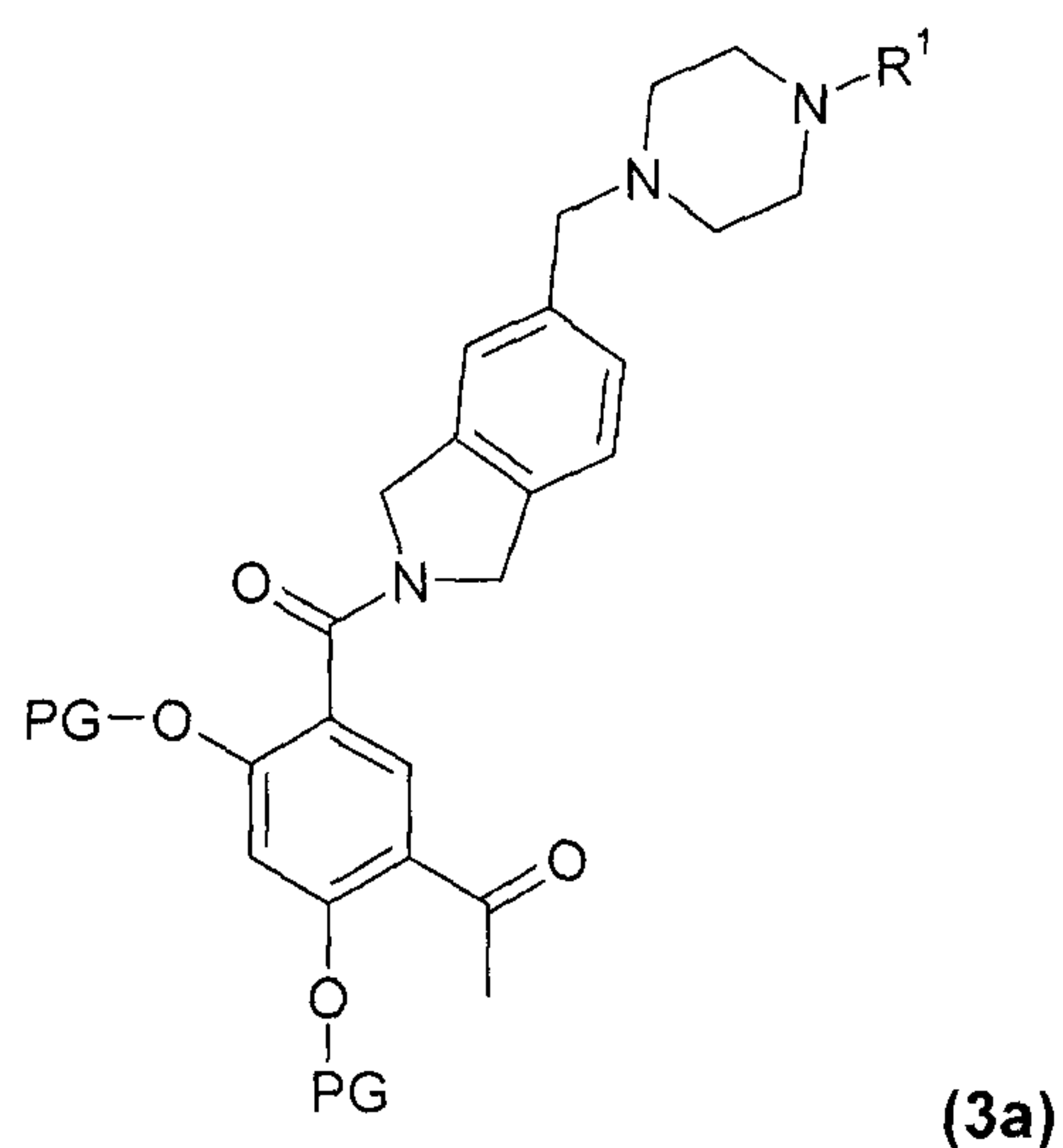
(a-i) the reaction of a compound of the formula (4), or an activated form or derivative thereof with a compound of the formula (5):



under amide forming conditions; or

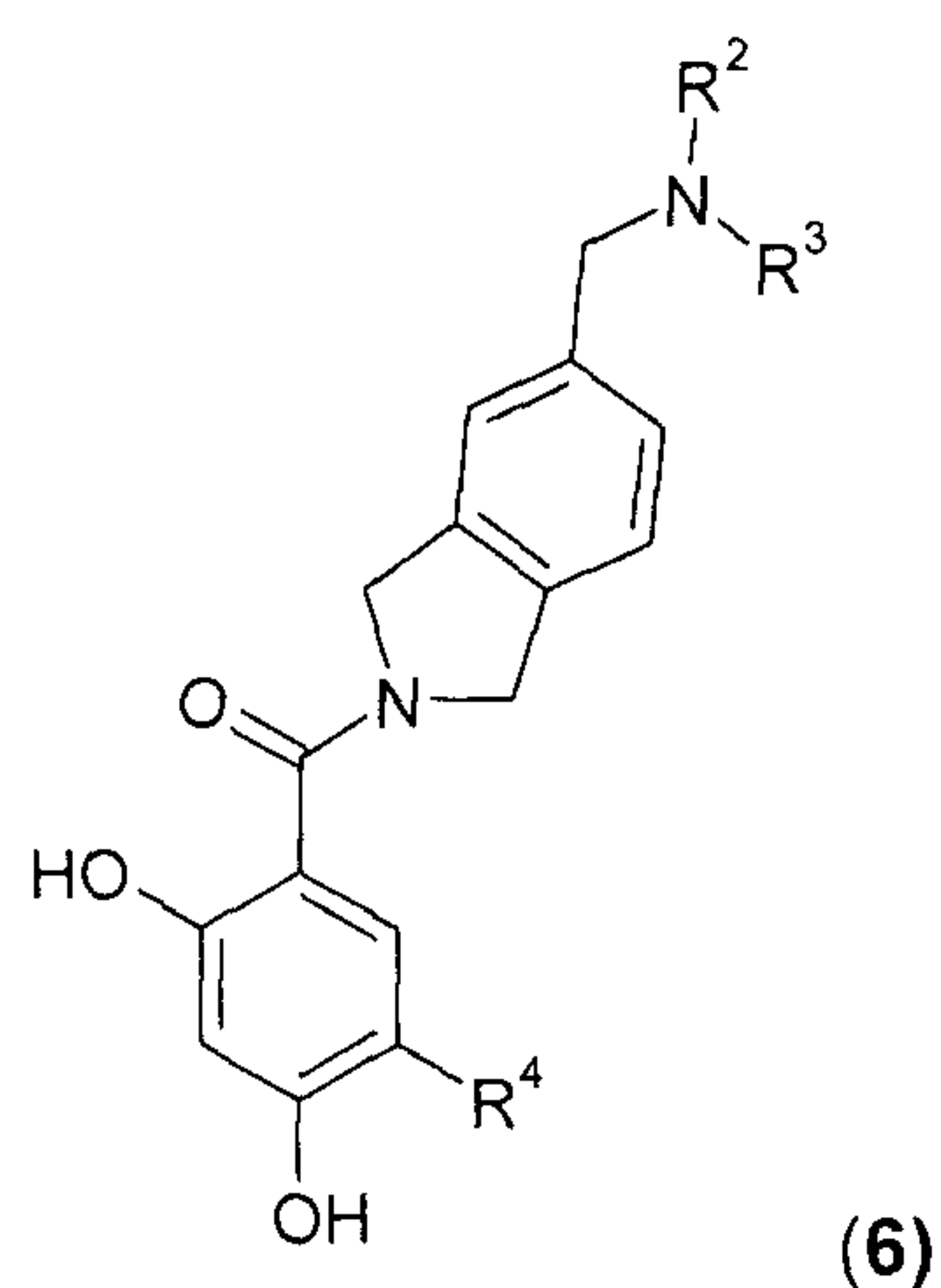
the reaction of a compound of the formula (3a):

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with a Wittig reagent or other reagent suitable for converting the group –
C(=O)-CH₃ into a group -C(=CH₂)-CH₃

6. A process for the preparation of a compound of the formula **(6)**:

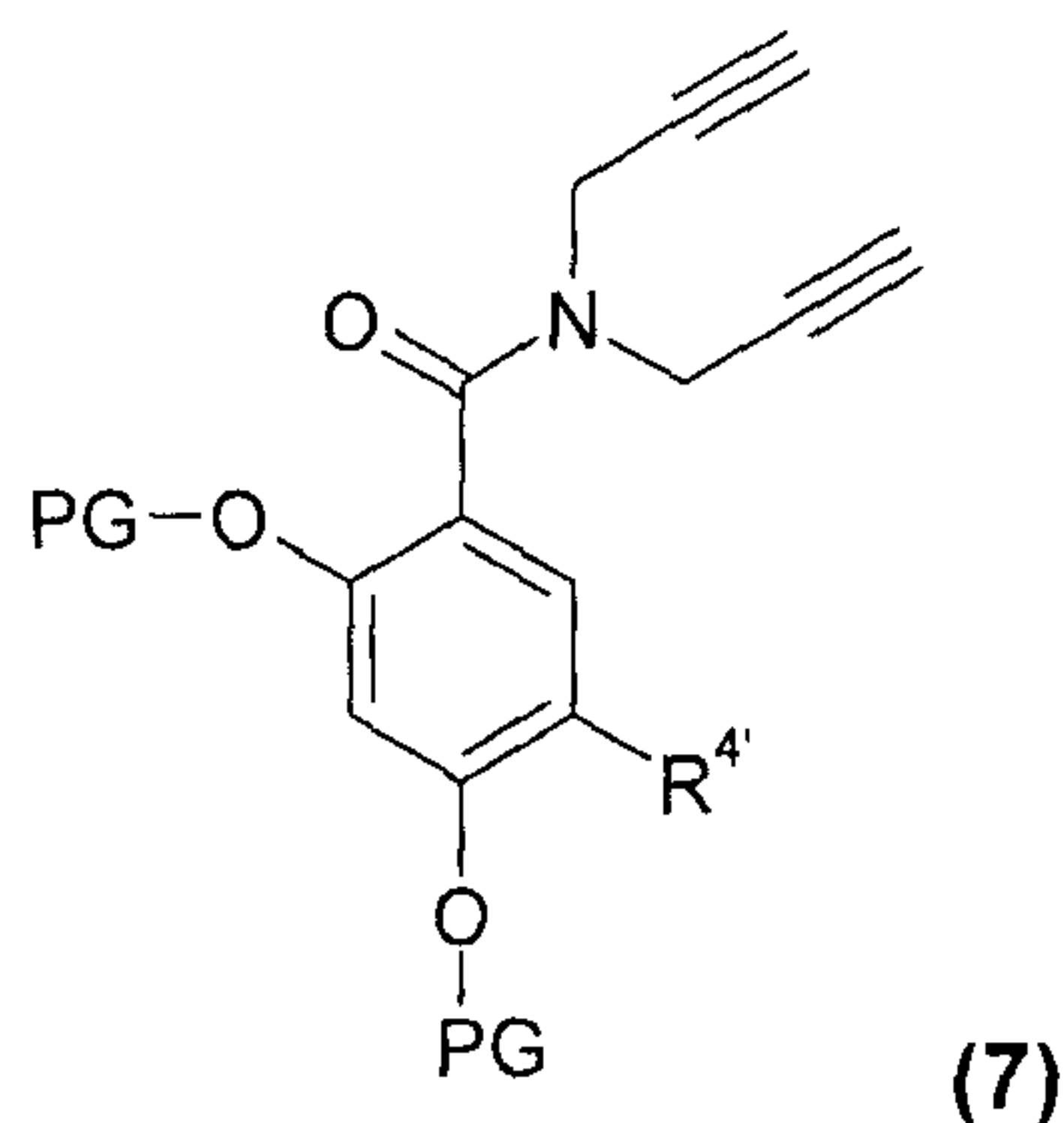


wherein R² and R³ are the same or different and each is C₁₋₄ alkyl or NR²R³ forms a 4 to 7 membered saturated heterocyclic ring optionally containing a further heteroatom selected from O, N and S and optionally substituted by one or two C₁₋₄ alkyl groups; and R⁴ is selected from hydrogen, halogen, C₁₋₅ alkyl and C₃₋₄ cycloalkyl groups;

which process comprises:

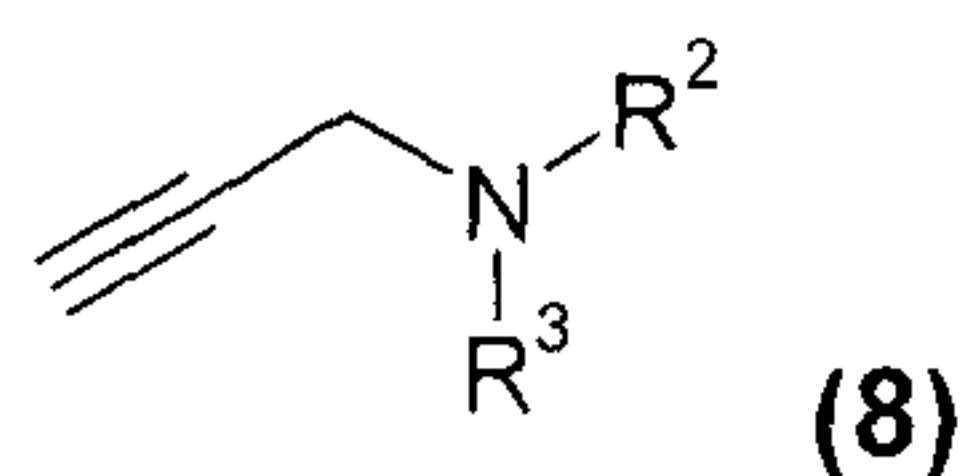
- (a-ii) the reaction of a compound of the formula **(7)**:

145

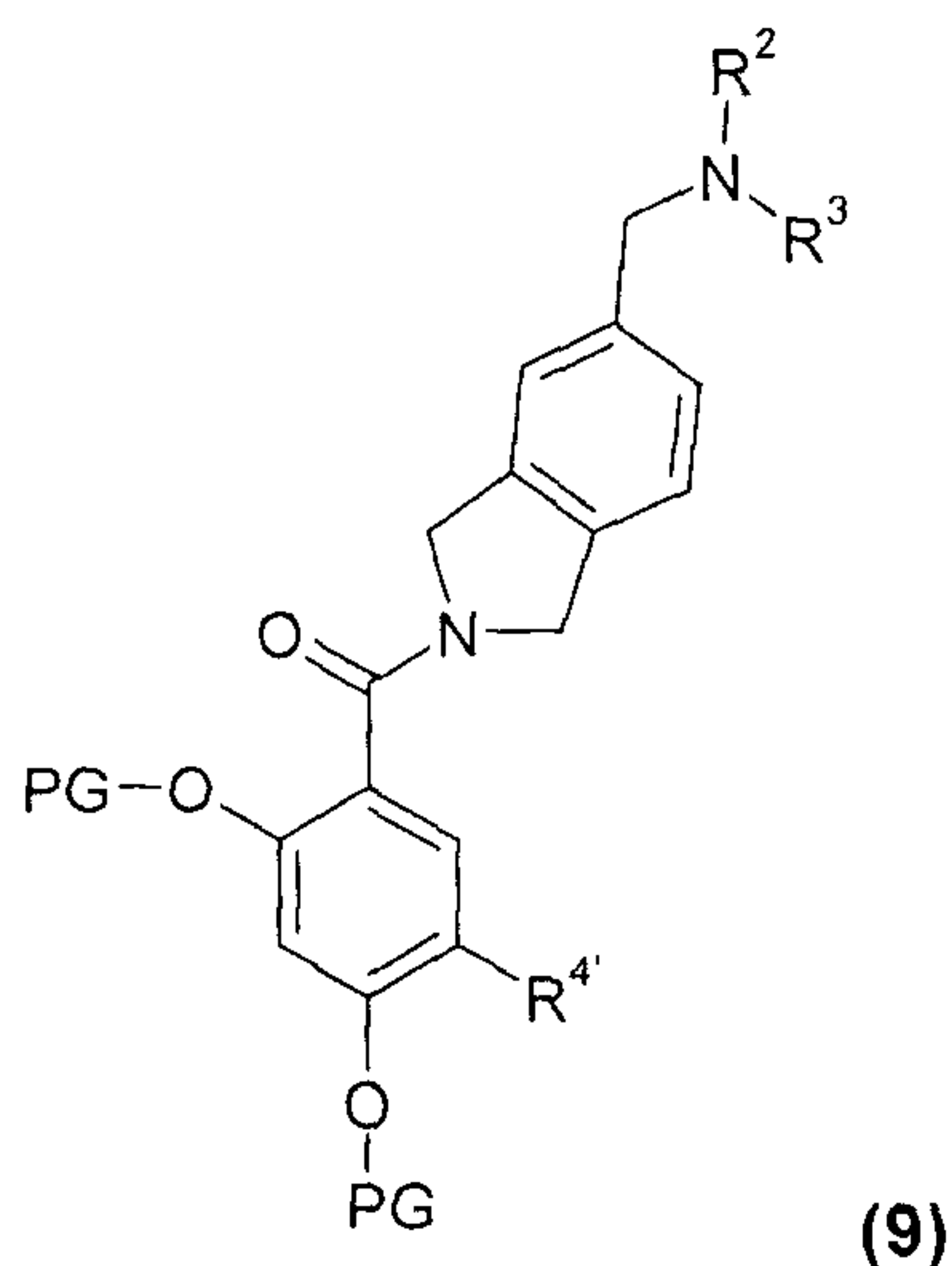


wherein PG is a protecting group removable under hydrogenation conditions and R^{4'} is selected from hydrogen, halogen, C₁₋₅ alkyl, C₂₋₅ alkenyl and C₃₋₄ cycloalkyl groups;

with a compound of the formula (8):



in the presence of a transition metal catalyst to give a compound of the formula (9);

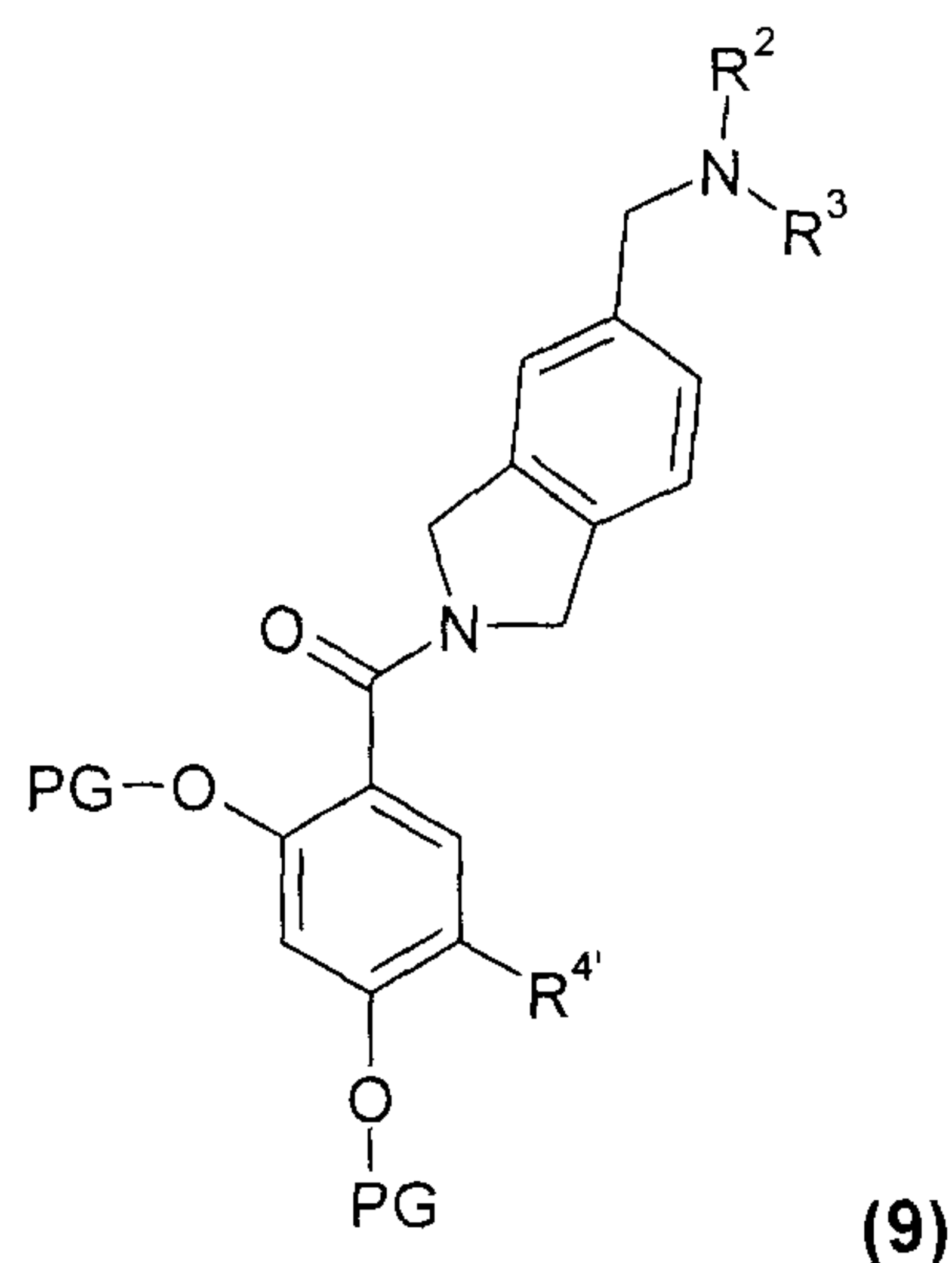


and

(b) subjecting the compound of formula (9) to catalytic hydrogenation to remove the protecting groups PG and, when R^{4'} is C₂₋₅ alkenyl, reduce the group R^{4'} to C₂₋₅ alkyl; and thereafter, where the compound of formula (6) is prepared in the form of a free base, optionally converting the free base to an acid addition salt.

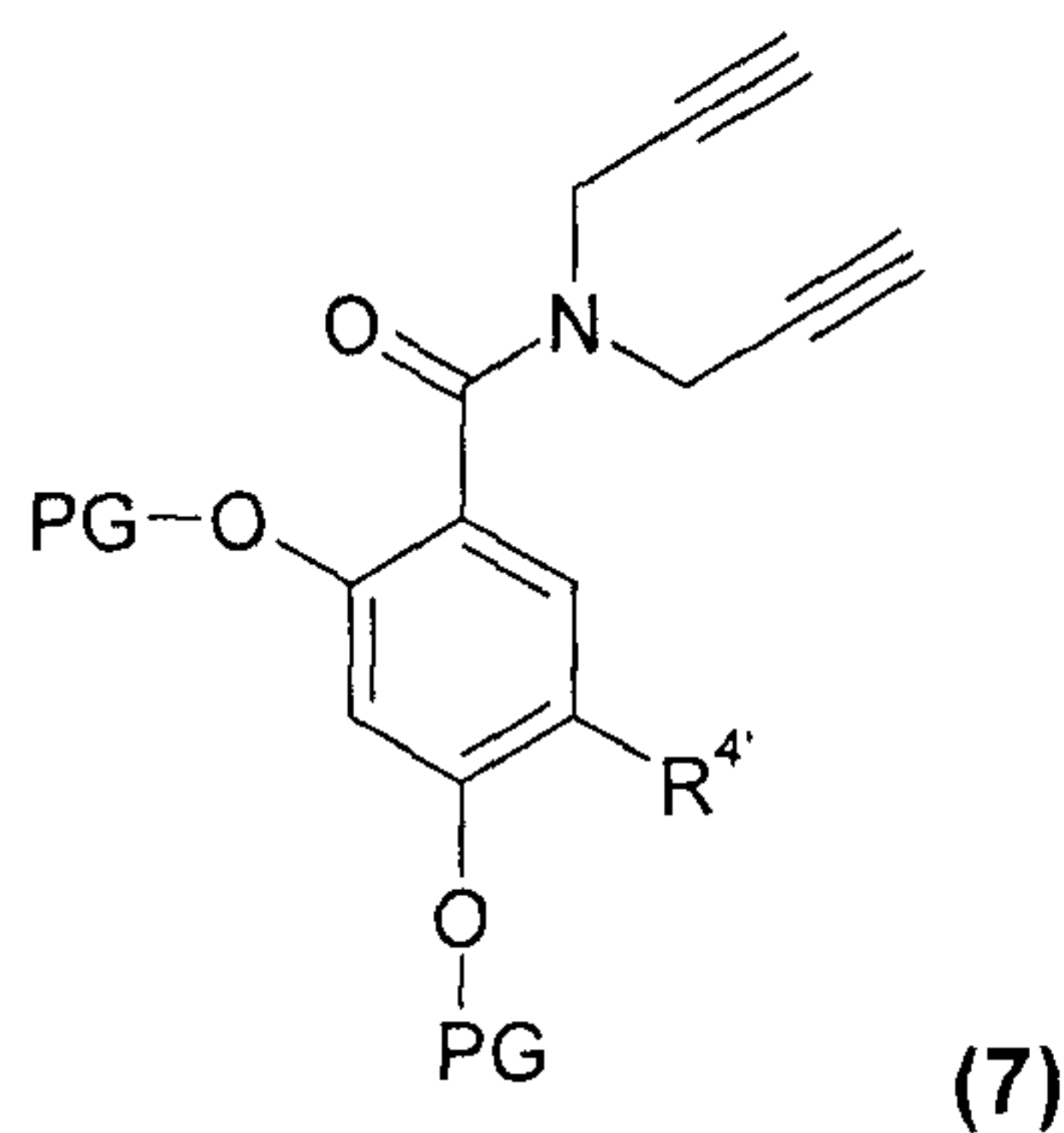
7. A process for the preparation of a compound of the formula (9):

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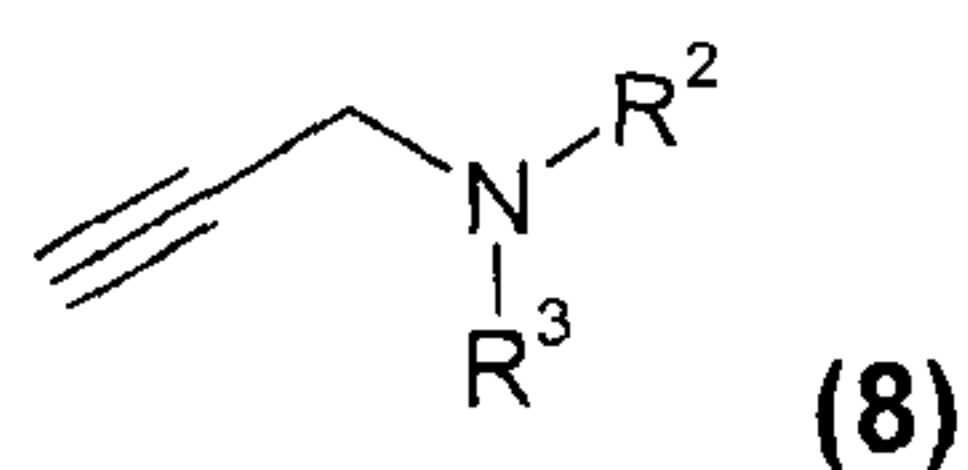
wherein R², R³, R^{4'} and PG are as defined in claim 6; which process comprises:

(a-ii) the reaction of a compound of the formula (7):



wherein PG is a protecting group removable under hydrogenation conditions and R^{4'} is selected from hydrogen, halogen, C₁₋₅ alkyl, C₂₋₅ alkenyl and C₃₋₄ cycloalkyl groups;

with a compound of the formula (8):



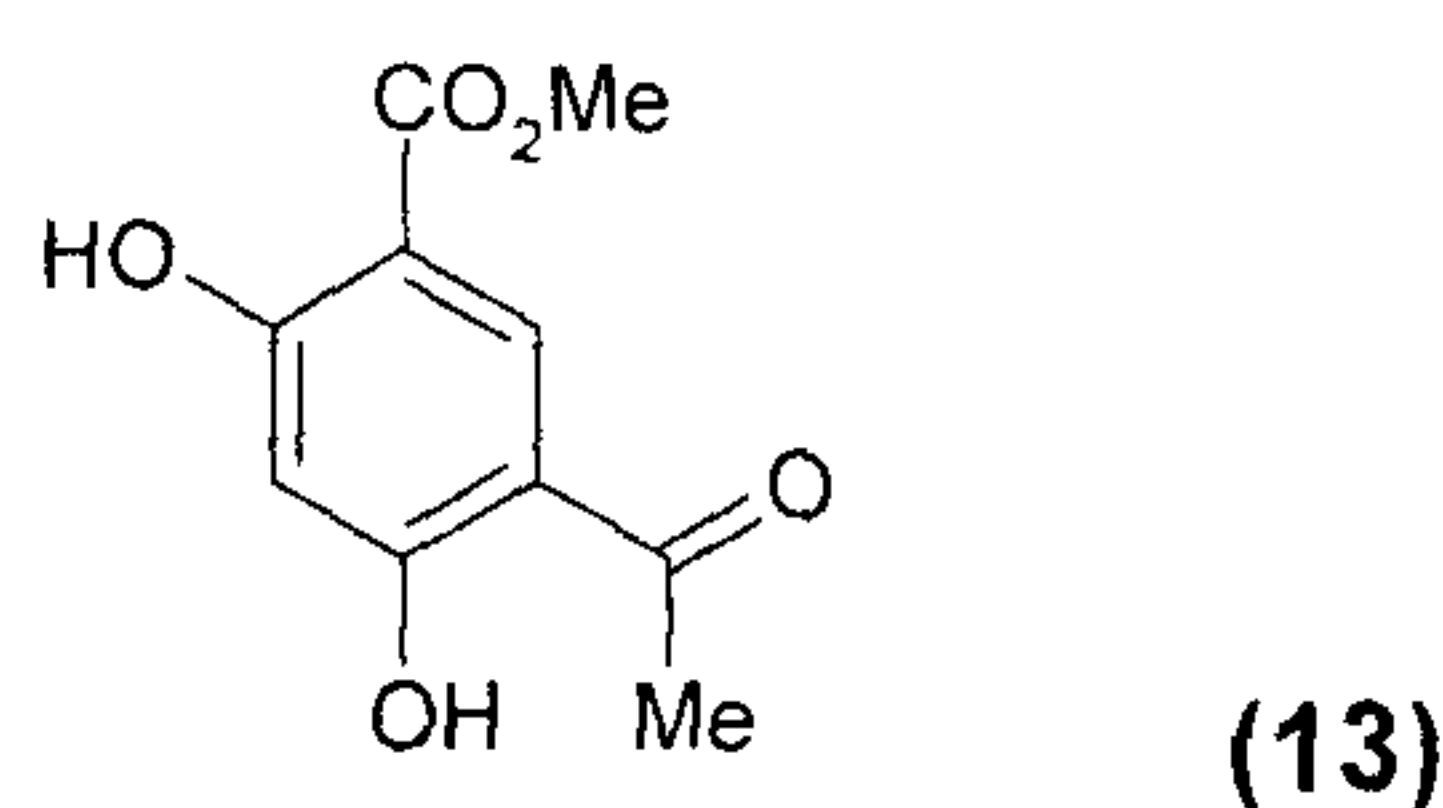
in the presence of a transition metal catalyst.

8. A chemical intermediate of the formula (3) as defined in claims 4 and 5, (7) as defined in claims 6 and 7, and (9) as defined in claims 6 and 7.
9. An acid addition salt according to any one of claims 1 to 3 for use:
 - i. as an inhibitor of Hsp90; or

- ii. in treating a disease or condition comprising or arising from abnormal cell growth in a mammal; or
- iii. in the treatment of a proliferative disorders selected from a carcinoma of the bladder, breast, colon, kidney, epidermis, liver, lung, oesophagus, gall bladder, ovary, pancreas, stomach, cervix, thyroid, prostate, gastrointestinal system, or skin; a hematopoietic tumour of lymphoid lineage; a hematopoietic tumour of myeloid lineage; thyroid follicular cancer; a tumour of mesenchymal origin; a tumour of the central or peripheral nervous system; melanoma; seminoma; teratocarcinoma; osteosarcoma; xeroderma pigmentosum; keratoacanthoma; or Kaposi's sarcoma; or
- iv. in the manufacture of a medicament for the prophylaxis or treatment of a fungal, protozoal or parasitic disease state or condition, other than a disease state or condition due to *Plasmodium falciparum*; or
- v. in the manufacture of a medicament for coadministration with an anti-fungal agent, anti-protozoal agent or anti-parasitic agent, to prevent, reduce or reverse the development of resistance to the anti-fungal agent, anti-protozoal agent or anti-parasitic agent; or
- vi. in the manufacture of a medicament for use with an anti-fungal agent, anti-protozoal agent or anti-parasitic agent to prevent or reduce development of resistance to an anti-fungal agent, anti-protozoal agent or anti-parasitic agent in a patient; or
- vii. in the manufacture of a medicament for use in the reduction or elimination of pain in a patient suffering from pain; or
- viii. in the manufacture of a medicament for the treatment of any one or more of nociception, somatic pain, visceral pain, acute pain, chronic pain, hyperalgesia, allodynia, post operative pain, pain due to hypersensitivity, headache, inflammatory pain, neurological pain, musculoskeletal pain, cancer related pain or vascular pain; or
- ix. in the manufacture of a medicament for the prevention or reduction of neuronal damage in a patient suffering from stroke; or

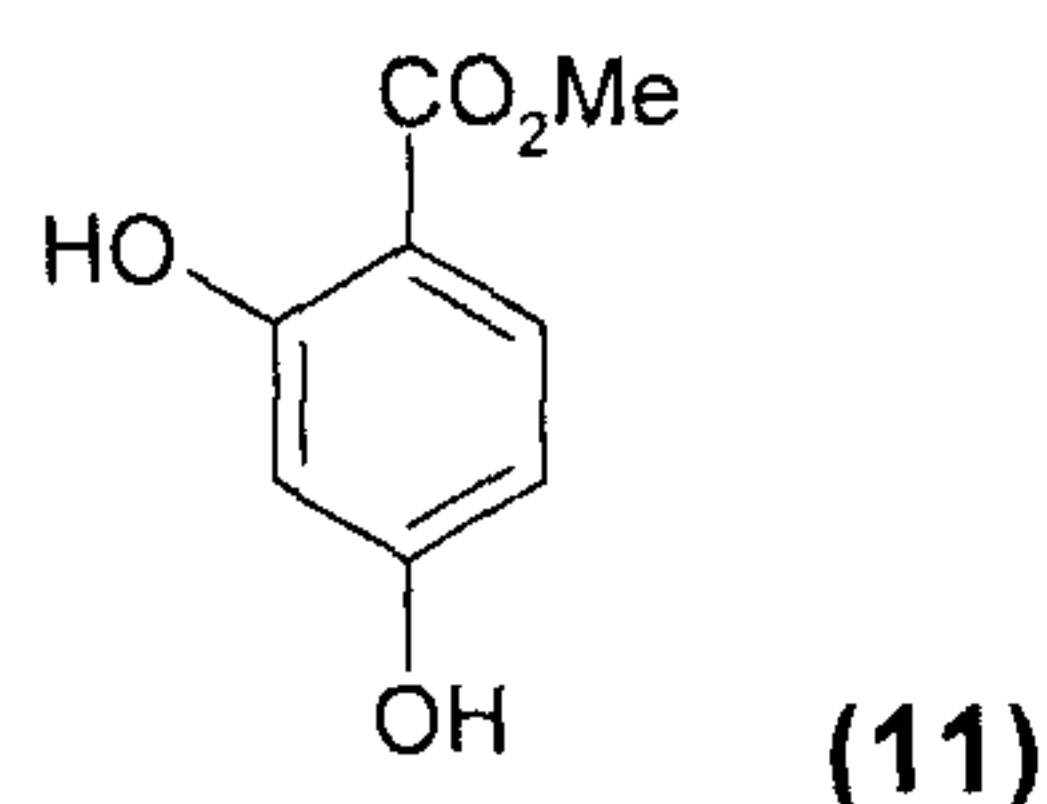
- x. in the manufacture of a medicament for the prevention or reduction of risk of stroke in patients at risk for stroke; or
 - xi. in the prophylaxis or treatment of a viral infection or viral disease; or
 - xii. in blocking or inhibiting viral replication in a host organism ; or
 - xiii. in: (i) sensitizing malignant cells to an anticancer drug; (ii) alleviating or reducing the incidence of resistance to an anticancer drug; (iii) reversing resistance to an anticancer drug; (iv) potentiating the activity of an anticancer drug; or (v) delaying or preventing the onset of resistance to an anticancer drug; or
 - xiv. in the treatment of a cancer characterized by the absence of drug resistance; or
 - xv. in the prophylaxis or treatment or alleviation or reduction of the incidence of a disease state or condition mediated by Hsp90 in a subject undergoing treatment with a therapeutic agent, wherein the disease state or condition mediated by Hsp90 is the development of resistance to the said therapeutic agent; or
 - xvi. in: (i) sensitizing malignant cells to an anti-cancer agent; (ii) alleviating or reducing the incidence of resistance to an anti-cancer agent; (iii) reversing resistance to an anti-cancer agent; (iv) potentiating the activity of an anti-cancer agent; or (v) delaying or preventing the onset of resistance to an anti-cancer agent; or
 - xvii. in the treatment of a cancer in a subject undergoing treatment with an anti-cancer agent characterized by the absence of drug resistance to the anti-cancer agent.
10. A pharmaceutical composition comprising an acid addition salt as defined in any one of claims 1 to 3, and a pharmaceutically acceptable carrier.
11. A process for the preparation of a compound of the formula **(13)**:

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which process comprises:

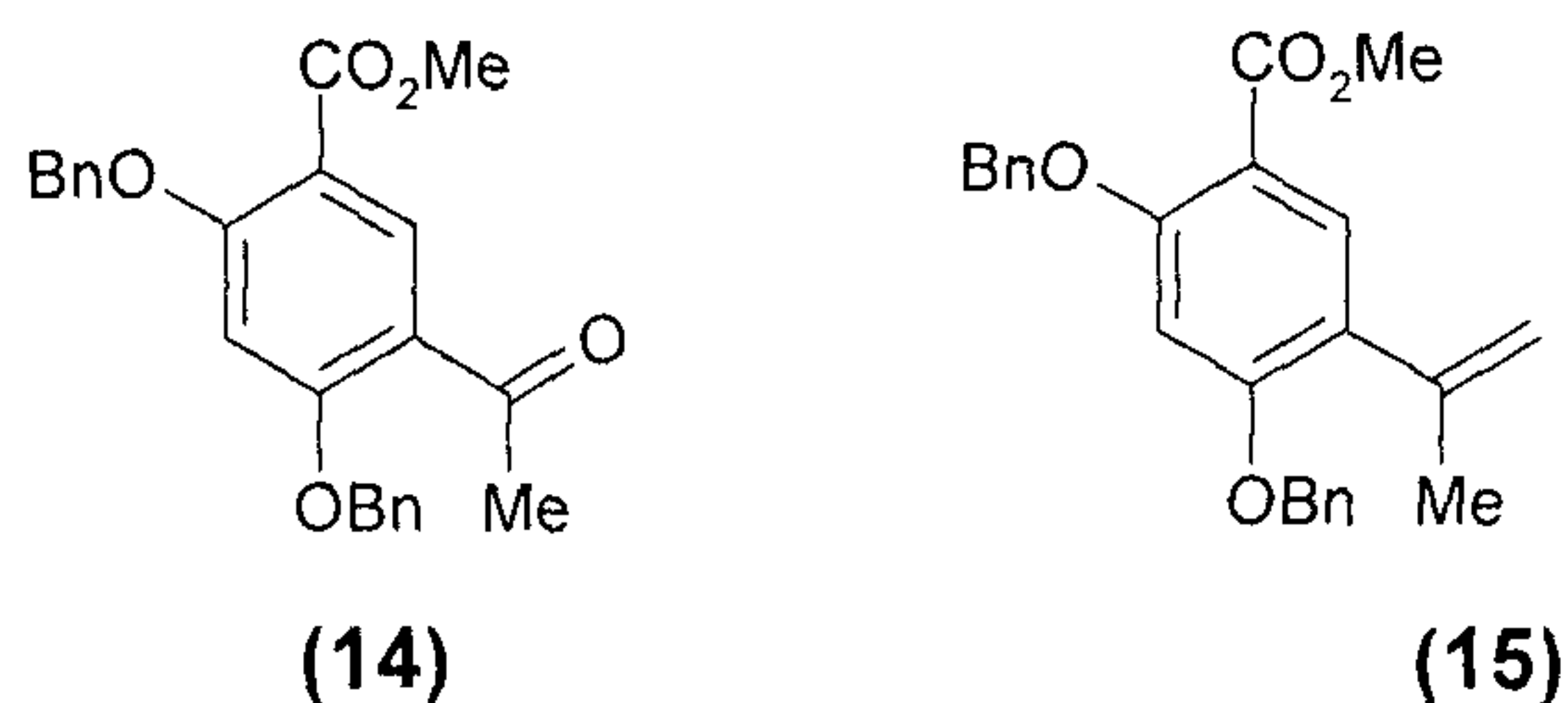
(i) the reaction of a compound of the formula **(11)**:



with (a) acetic anhydride in the presence of 4-dimethylaminopyridine, followed by (b) trifluoromethanesulphonic acid and optionally acetyl chloride (typically at room temperature); or

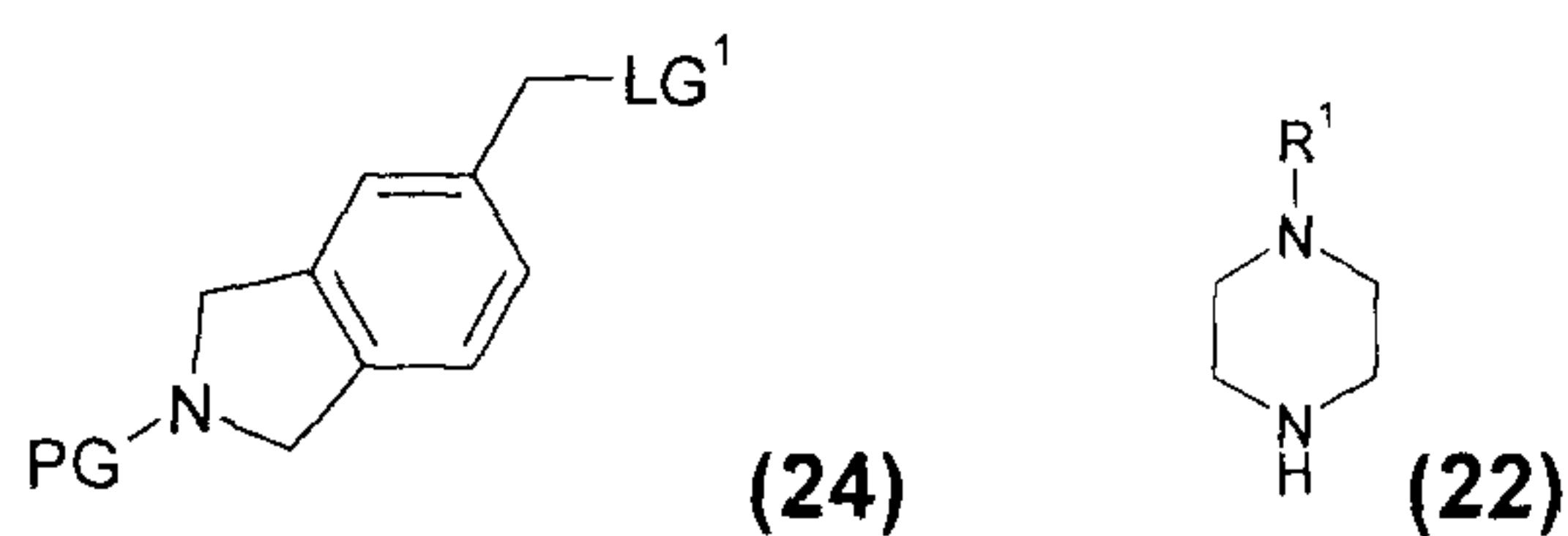
(ii) the reaction of a compound of the formula **(11)** with acetyl chloride in the presence of a cationic ion-exchange resin.

12. A process for the preparation of a compound of the formula **(15)**:



by reacting a compound of the formula **(14)** with a Wittig reagent MePPh_3Br in the presence of potassium tert-butoxide in THF.

13. A process for the preparation of a compound of the formula **(5)** as defined in claim 4, which process comprises the reaction of a compound of the formula **(24)**:



wherein PG is a protecting group and LG^1 is a leaving group, with a compound of the formula **(22)**.

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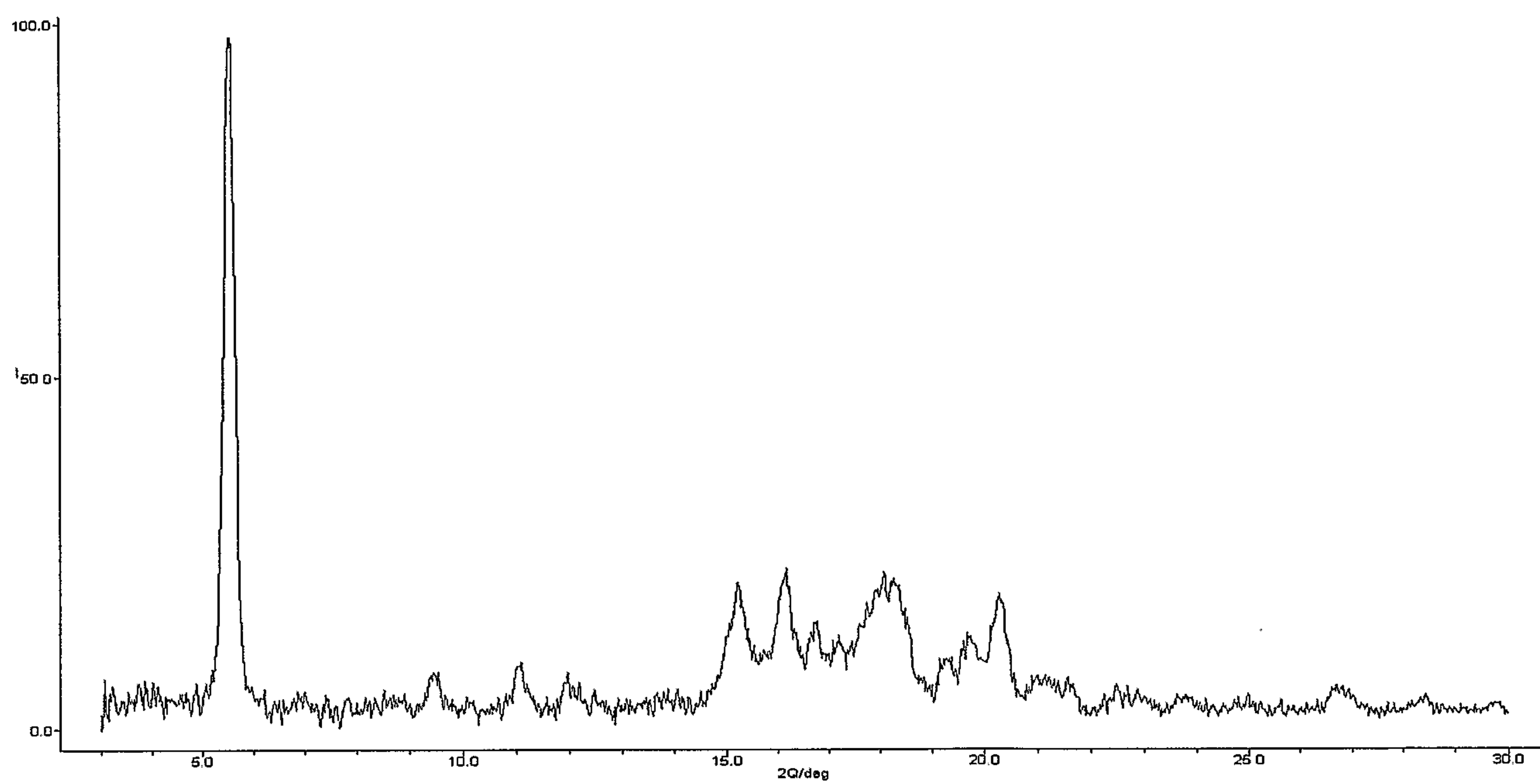


Figure 1

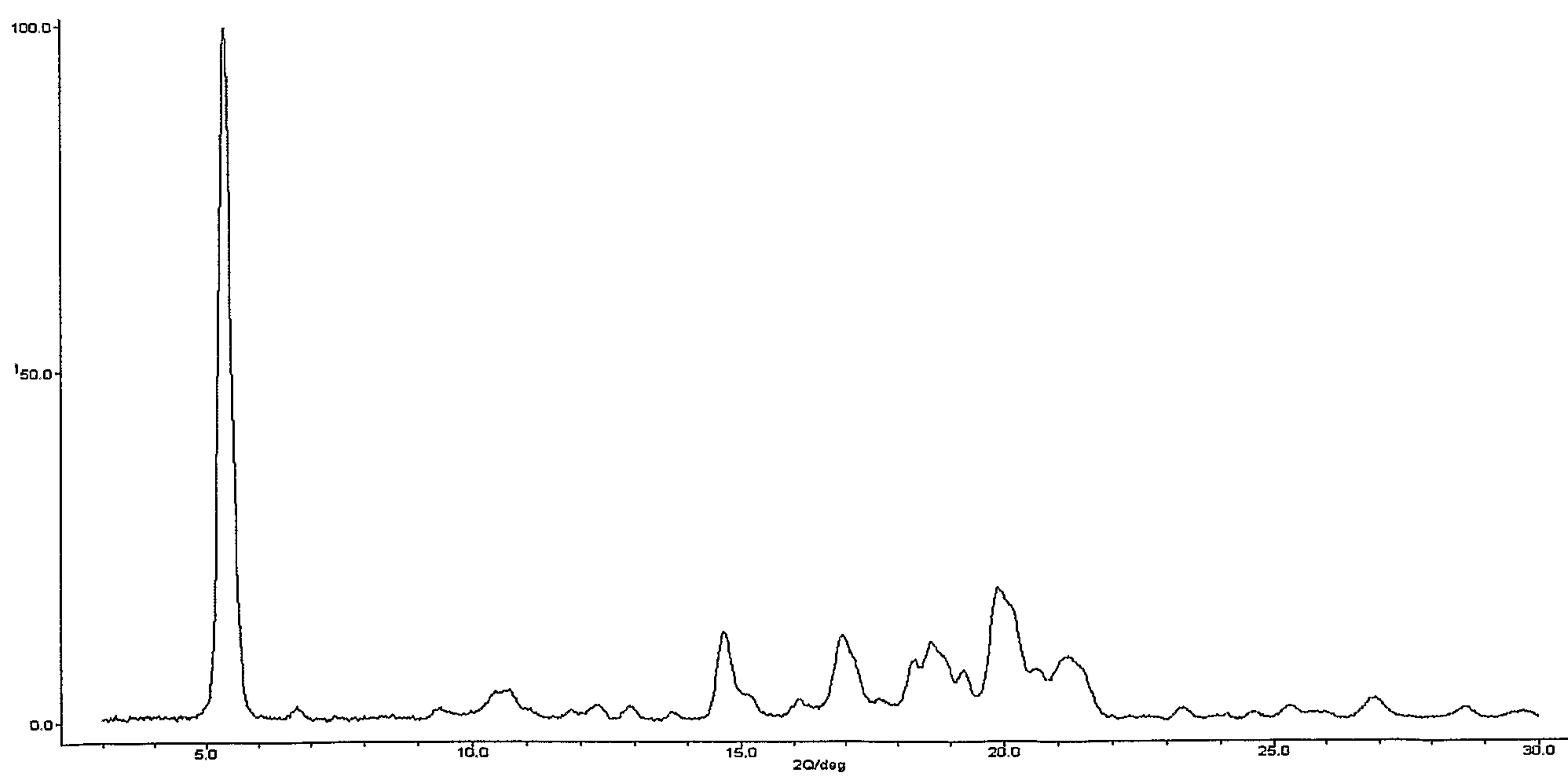


Figure 2

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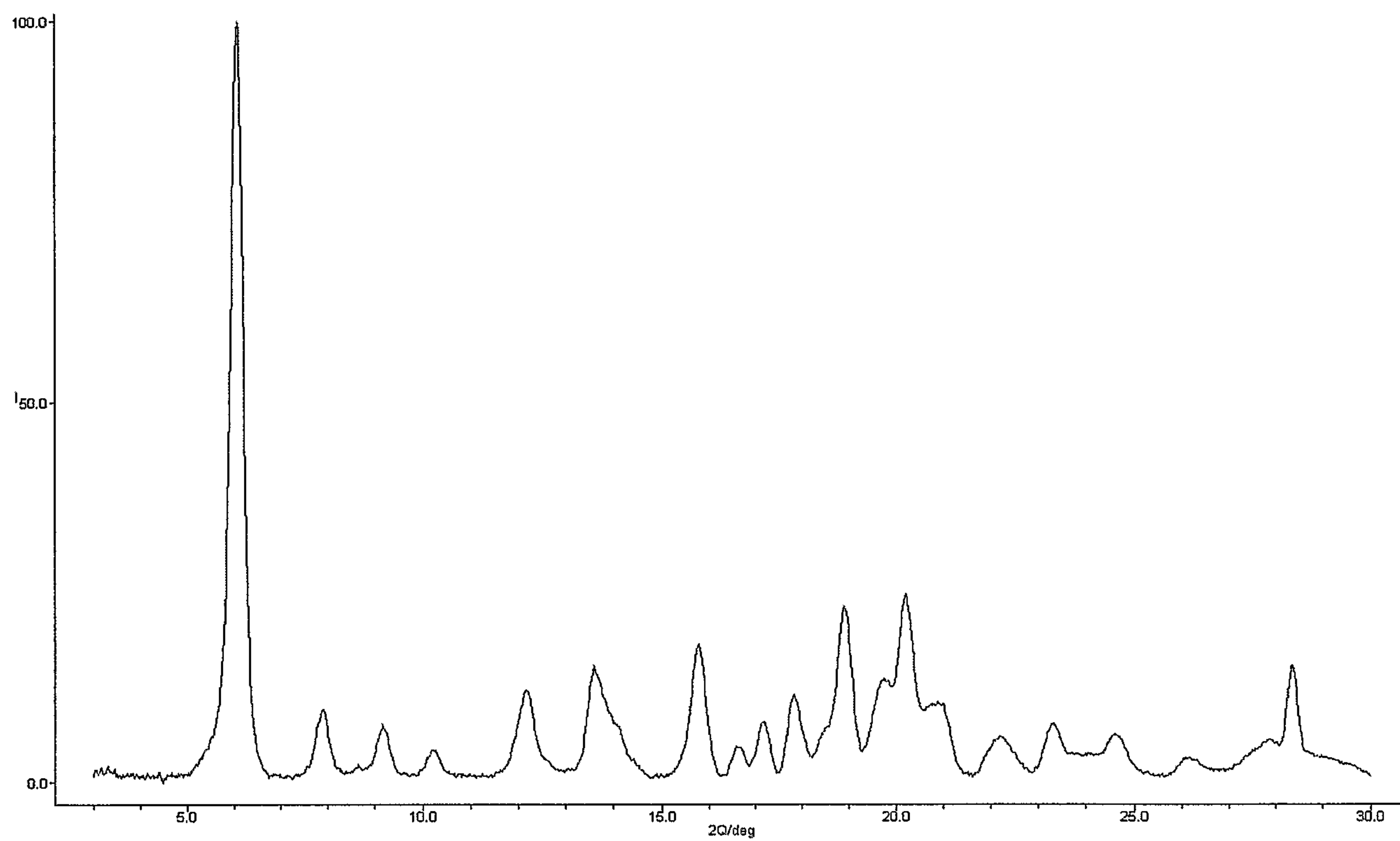


Figure 3

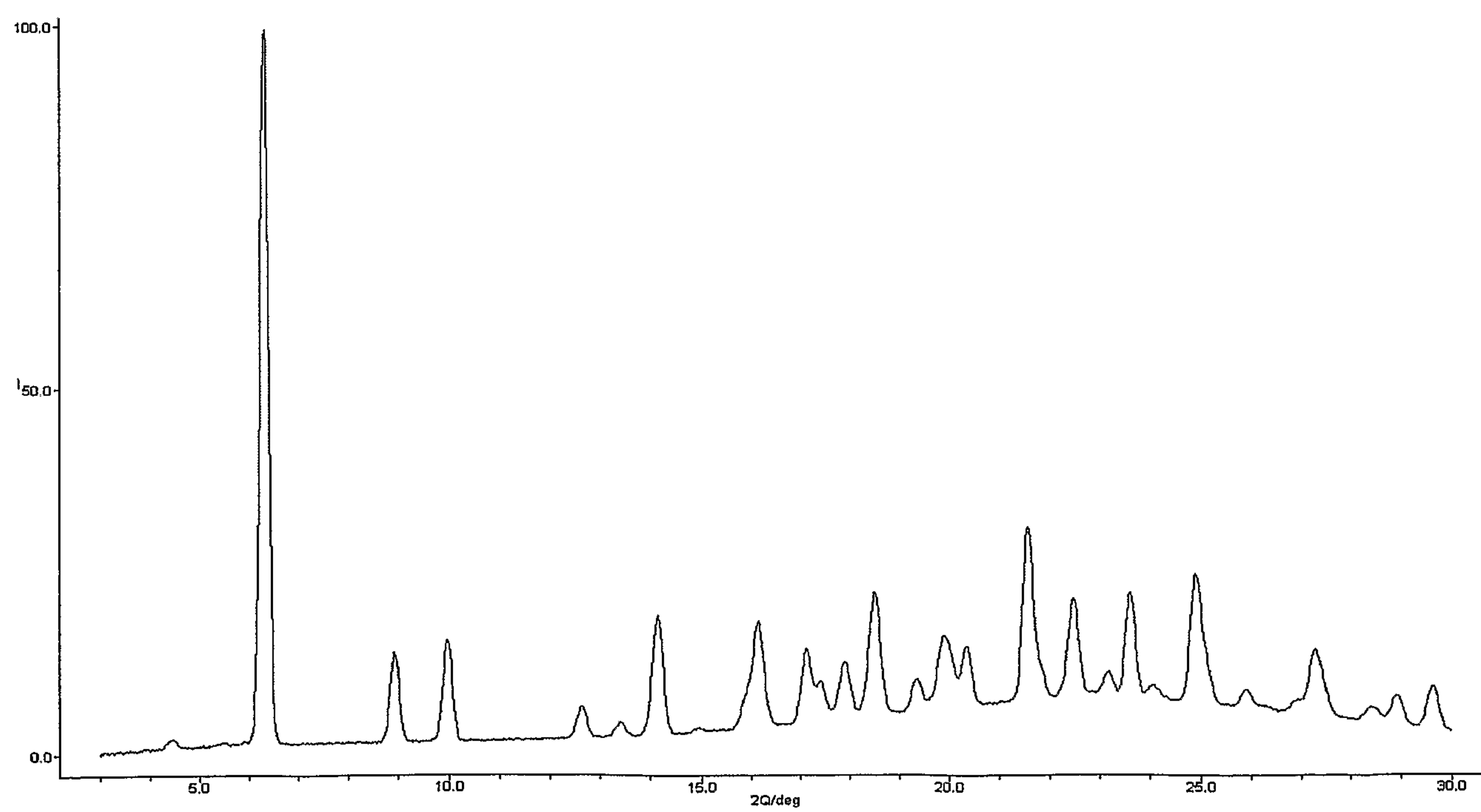


Figure 4

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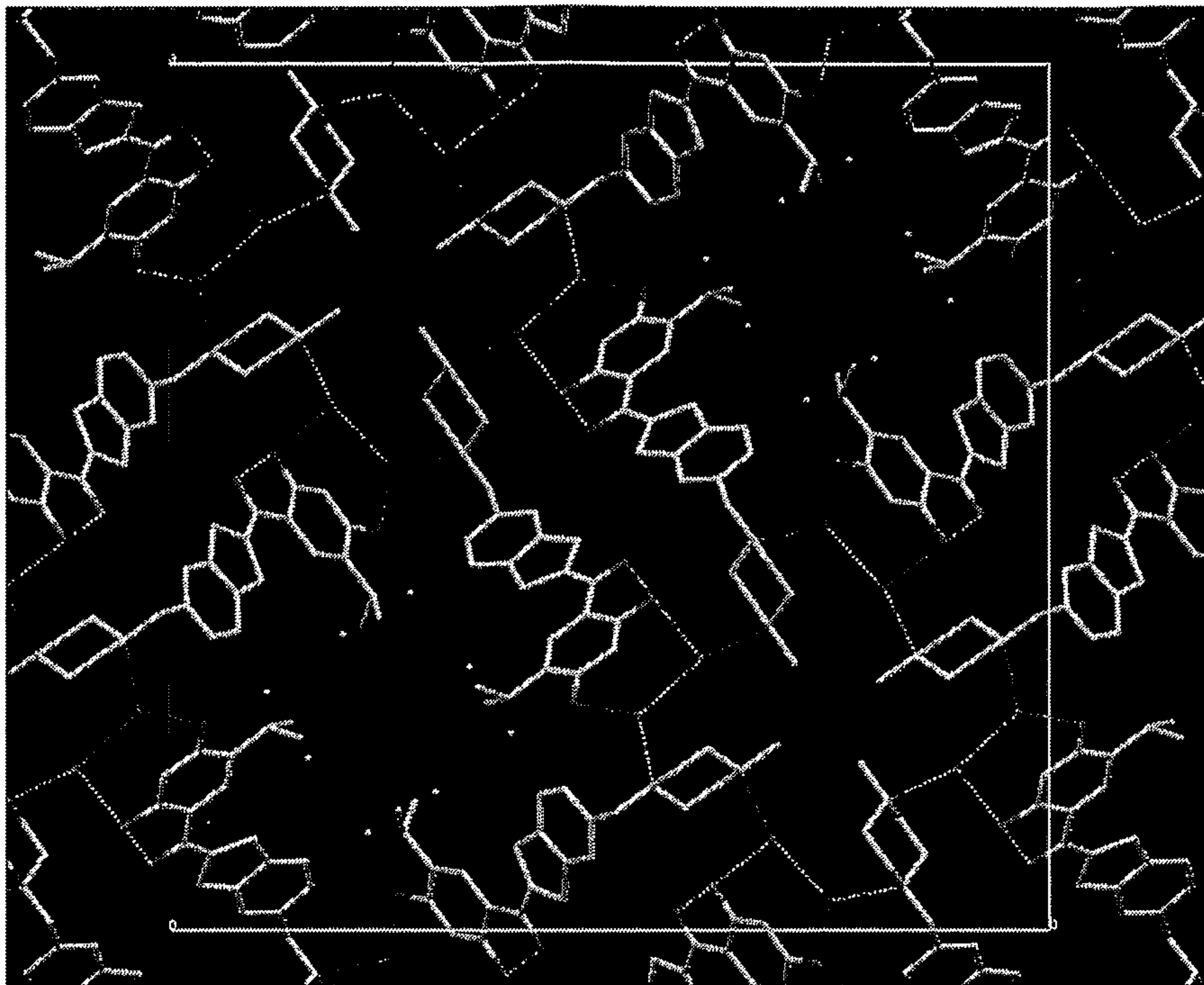


Figure 5

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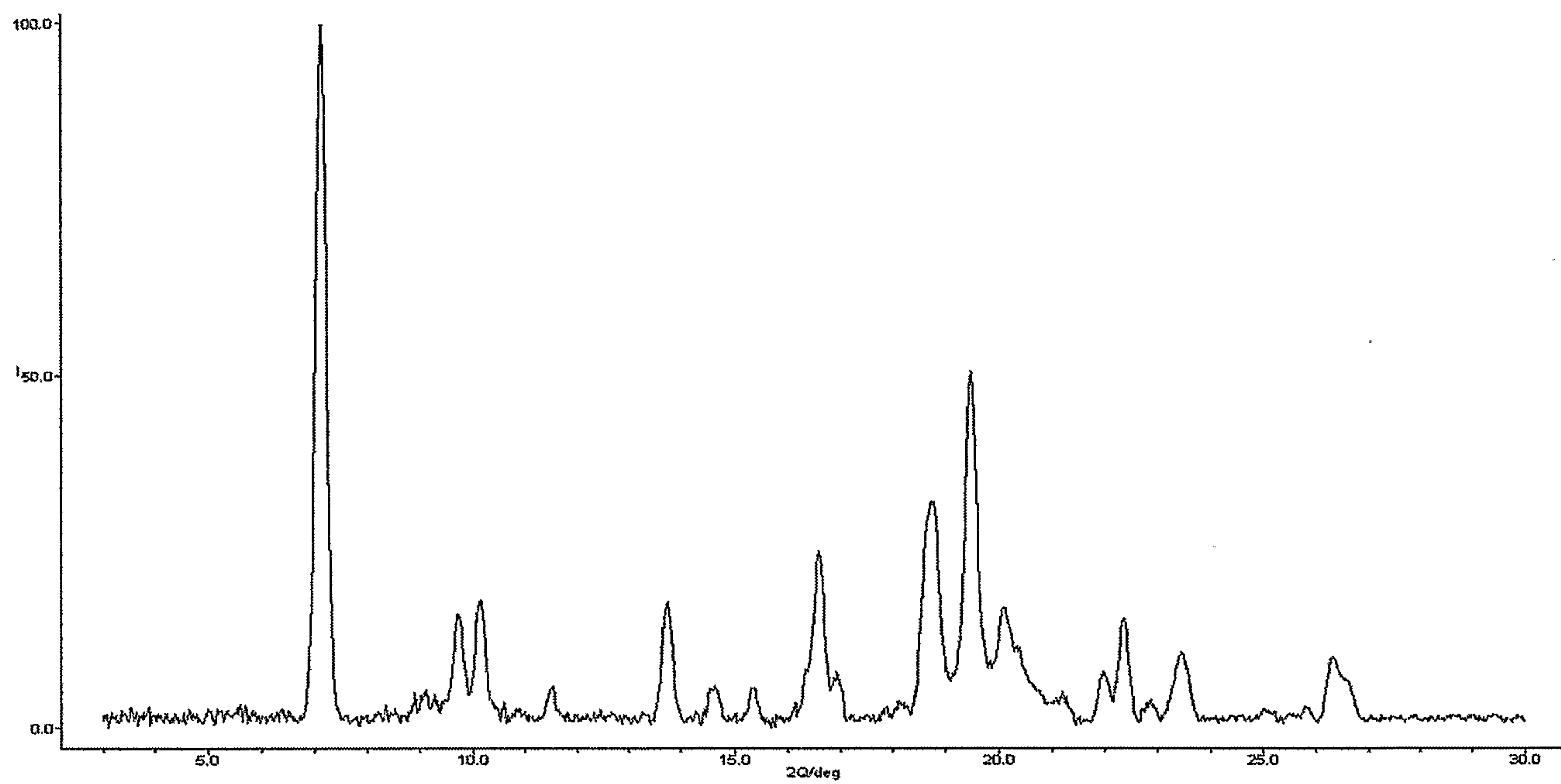


Figure 6

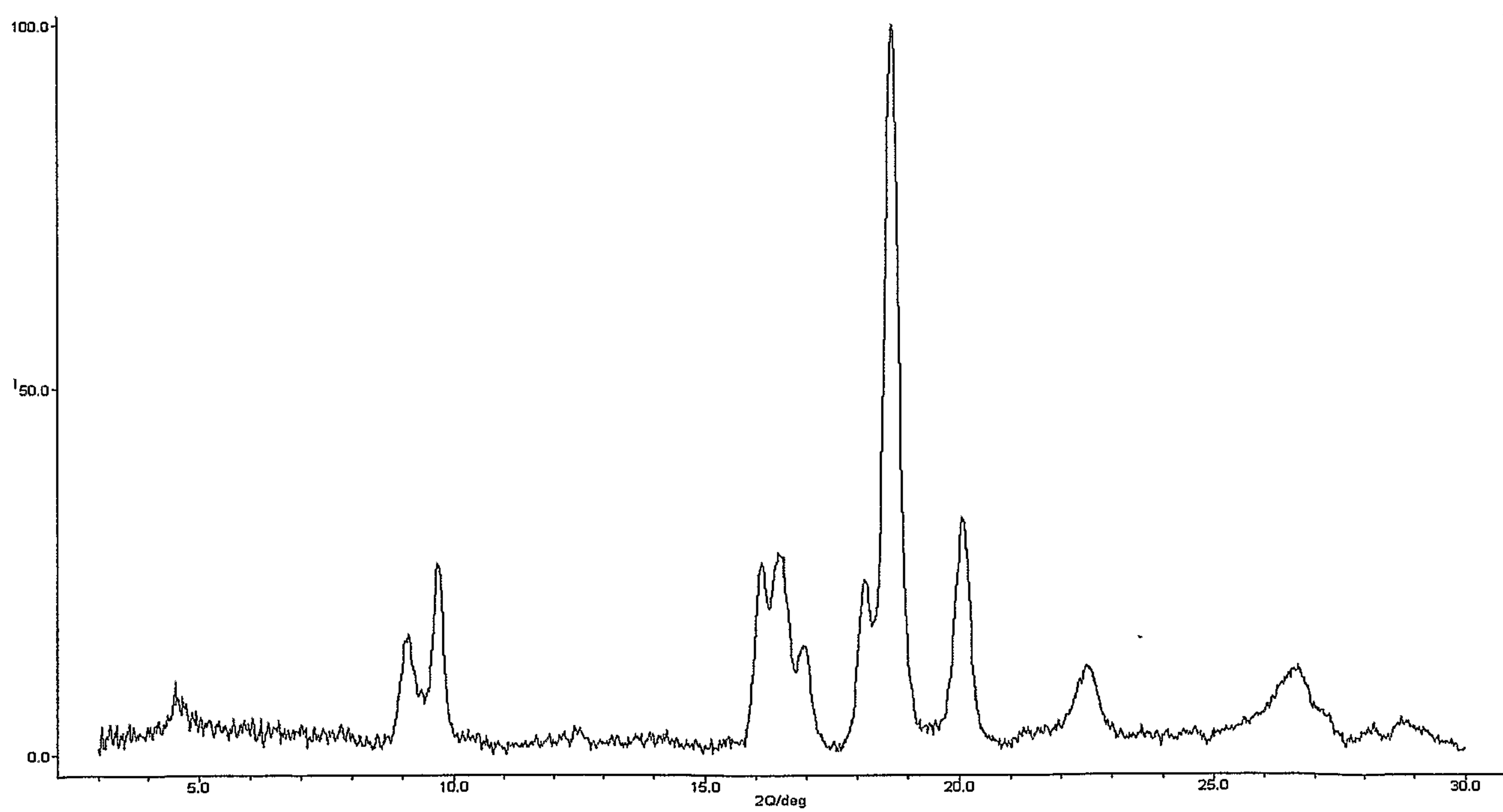


Figure 7

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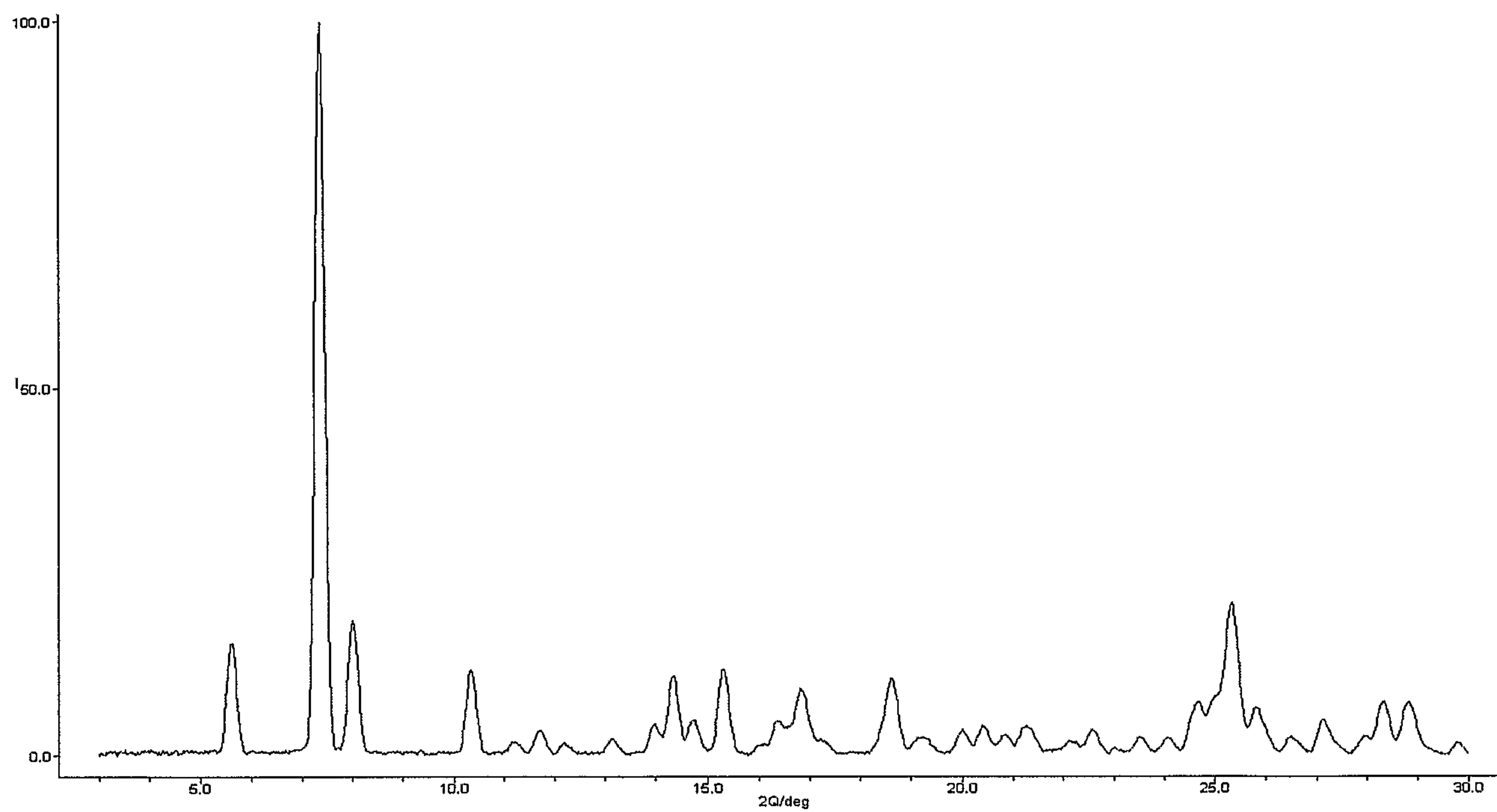


Figure 8

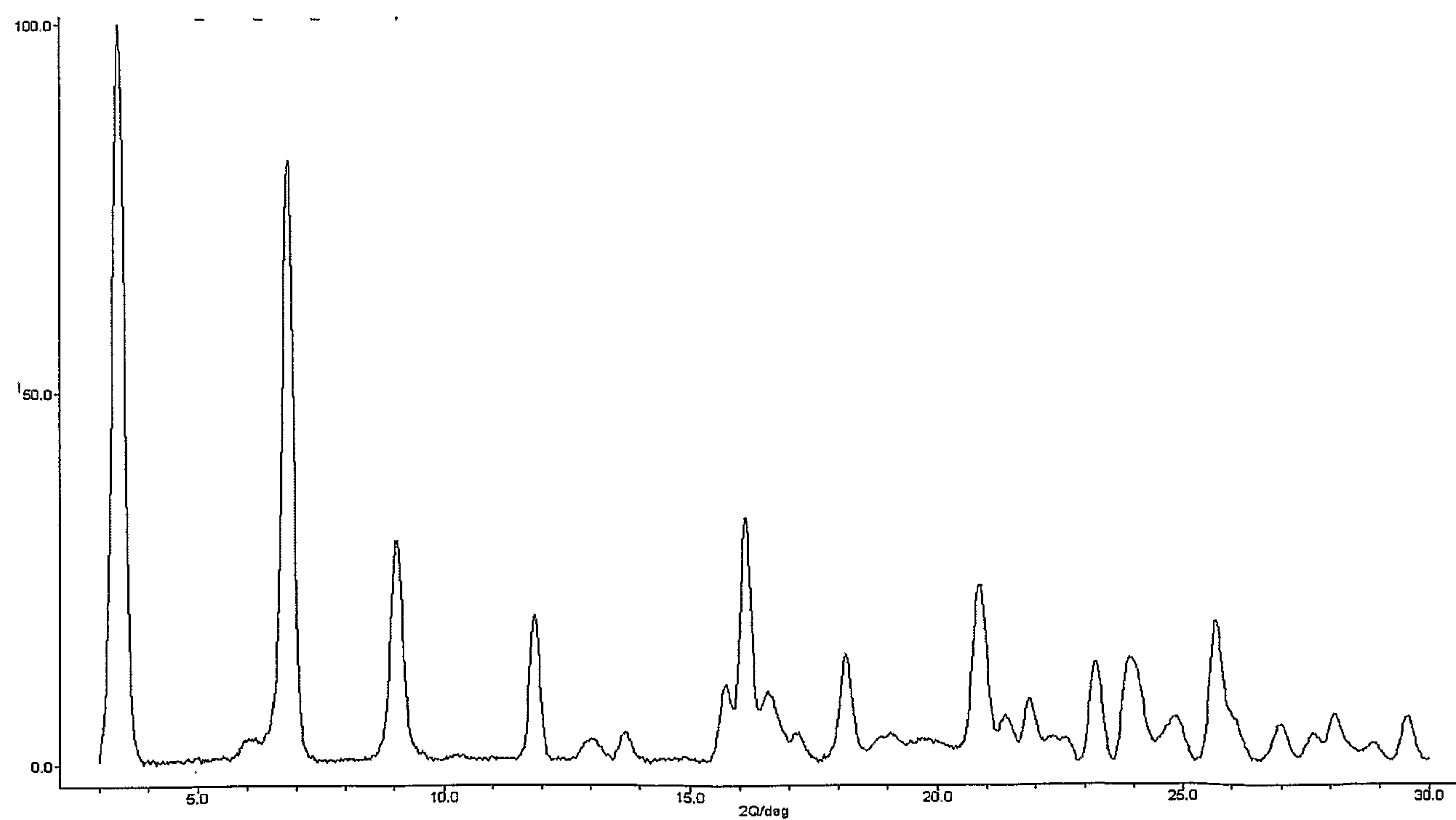


Figure 9

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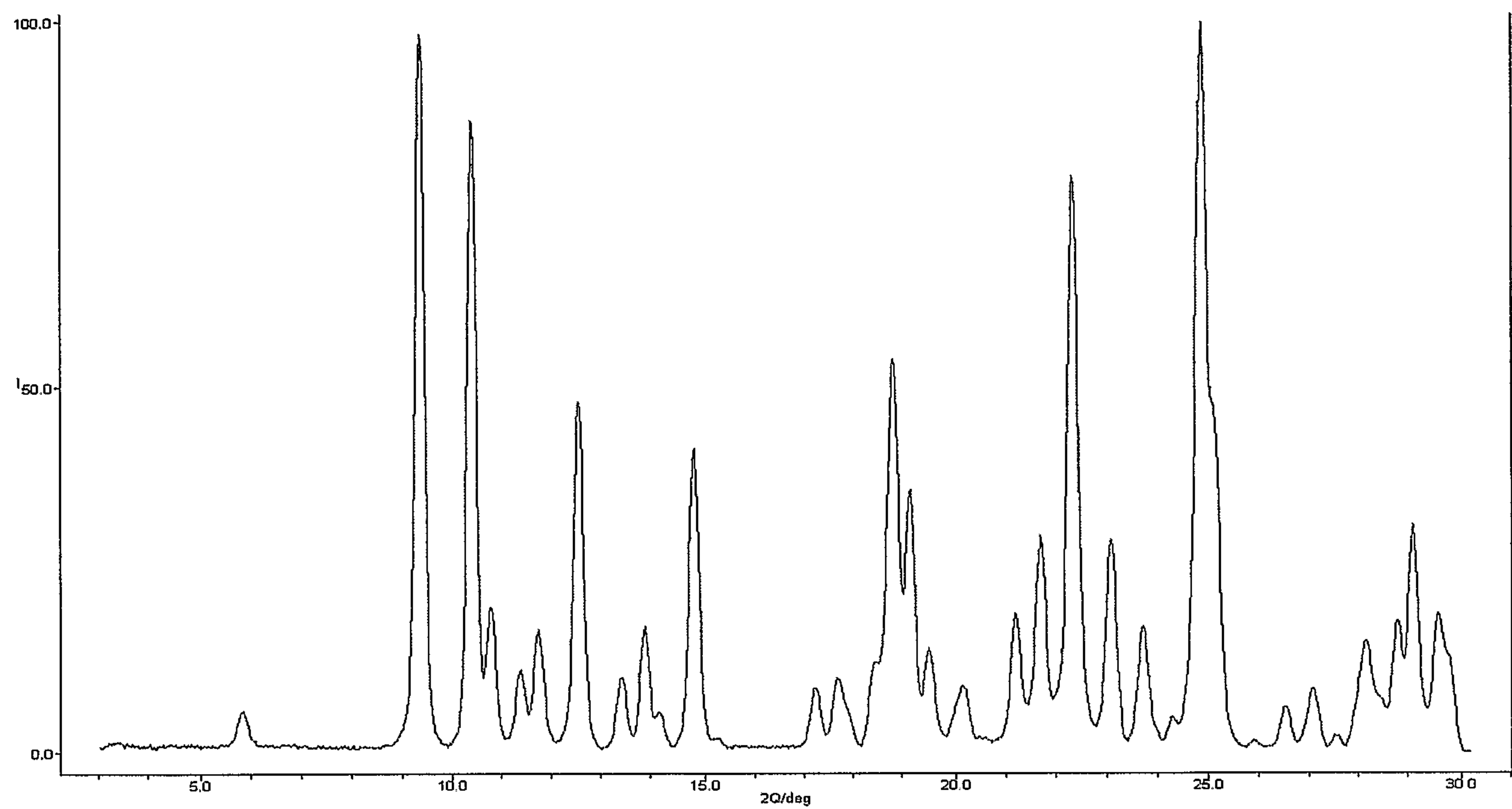


Figure 10

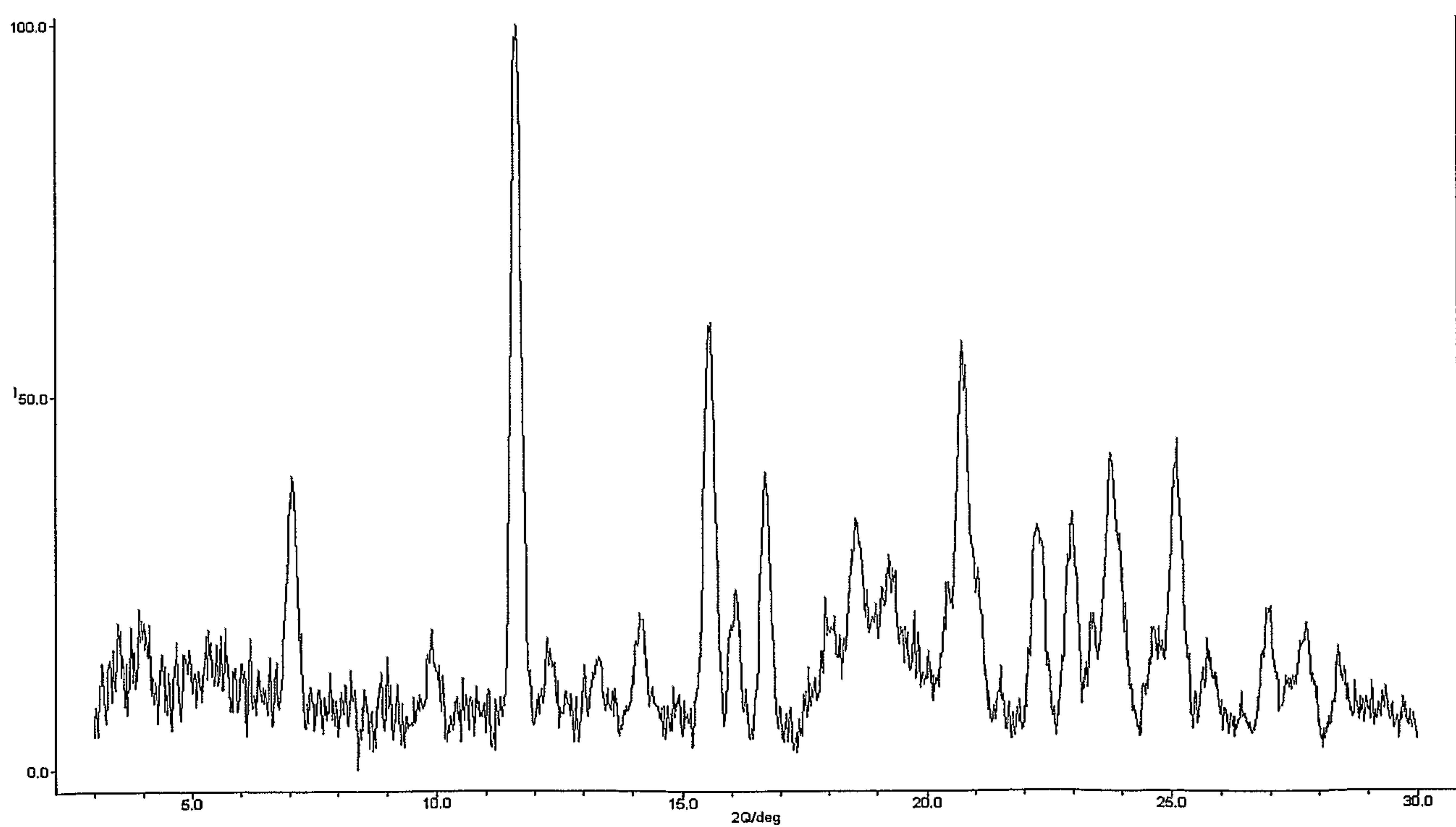


Figure 11

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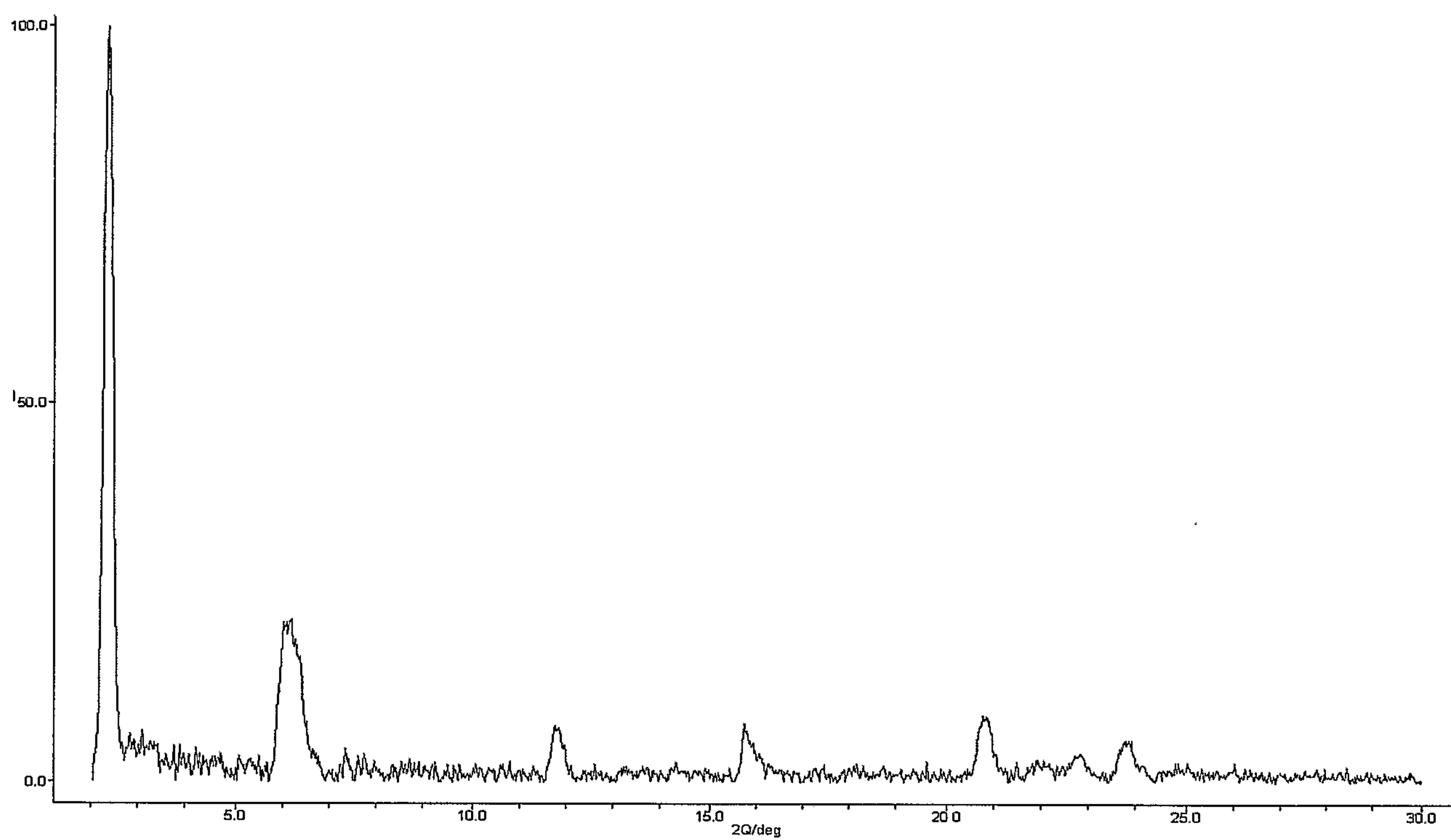


Figure 12

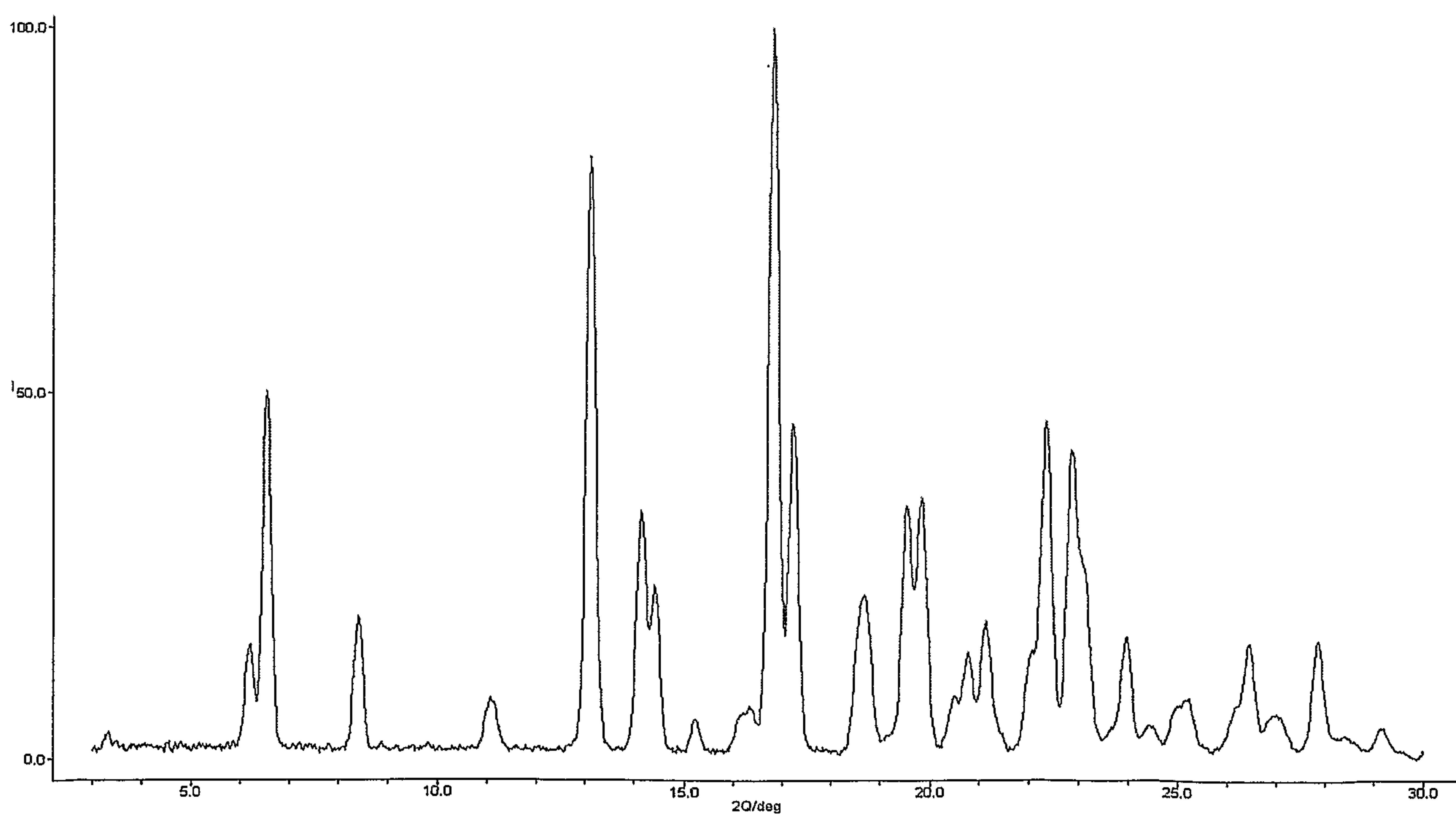


Figure 13

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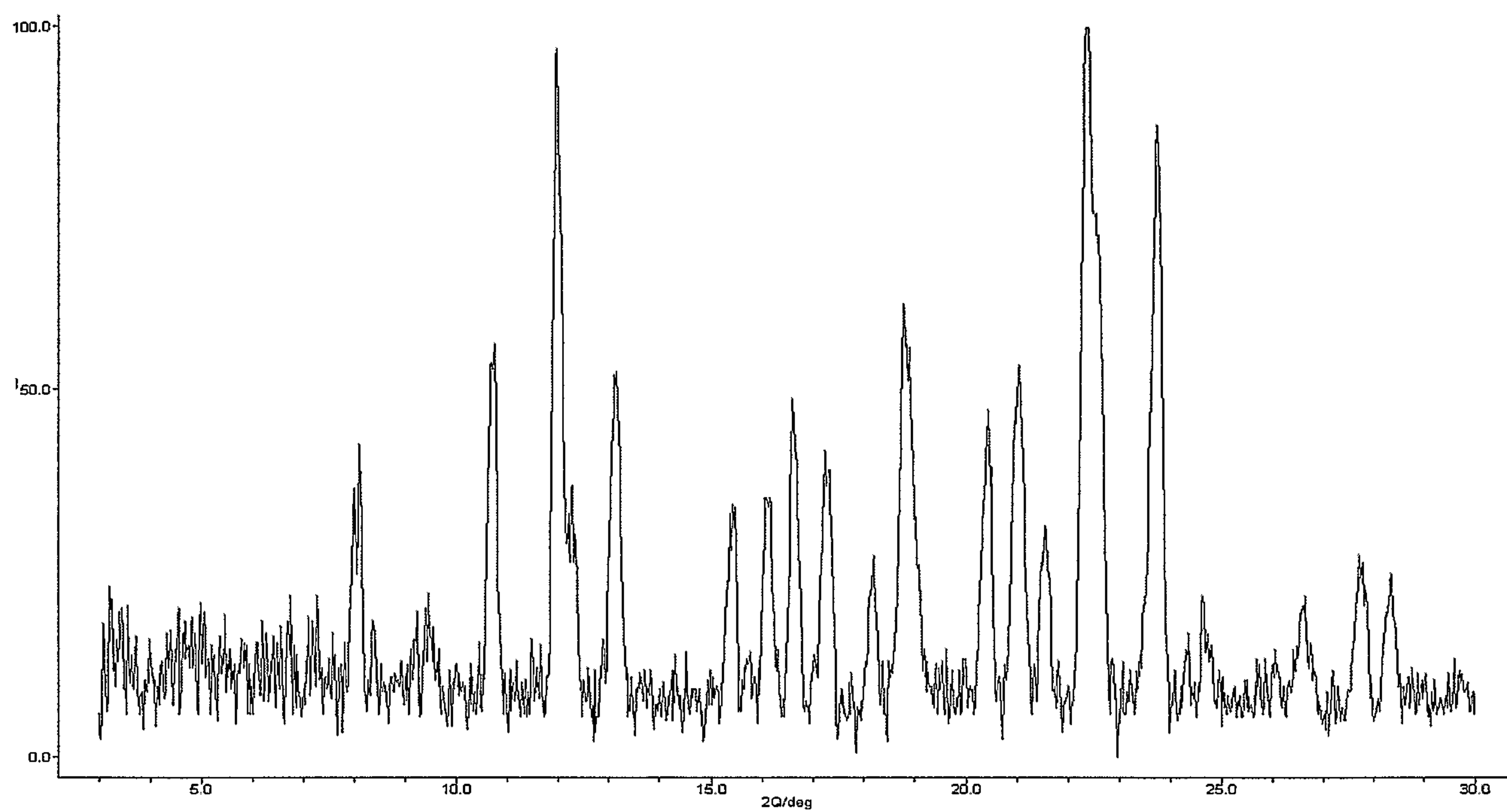


Figure 14

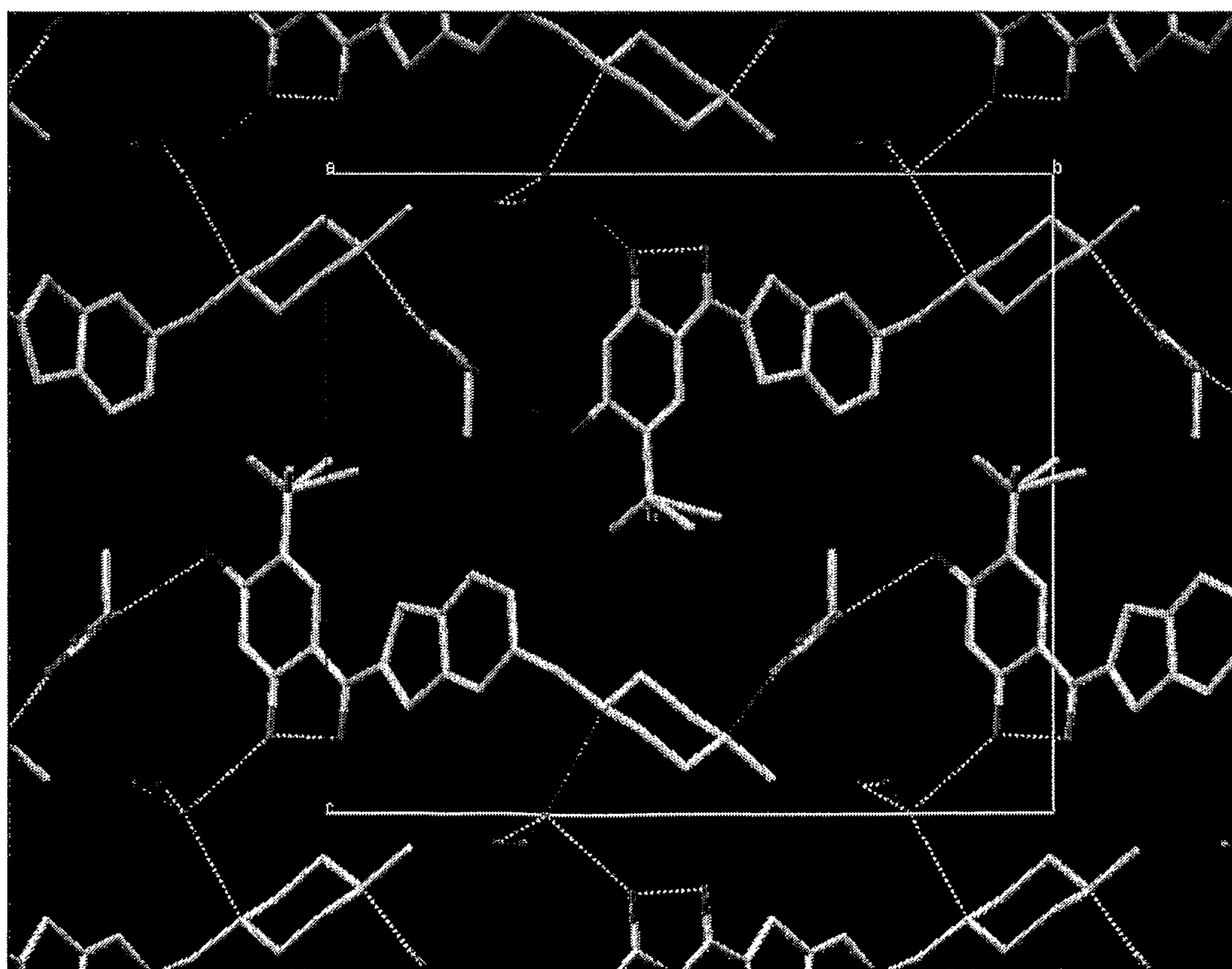


Figure 15

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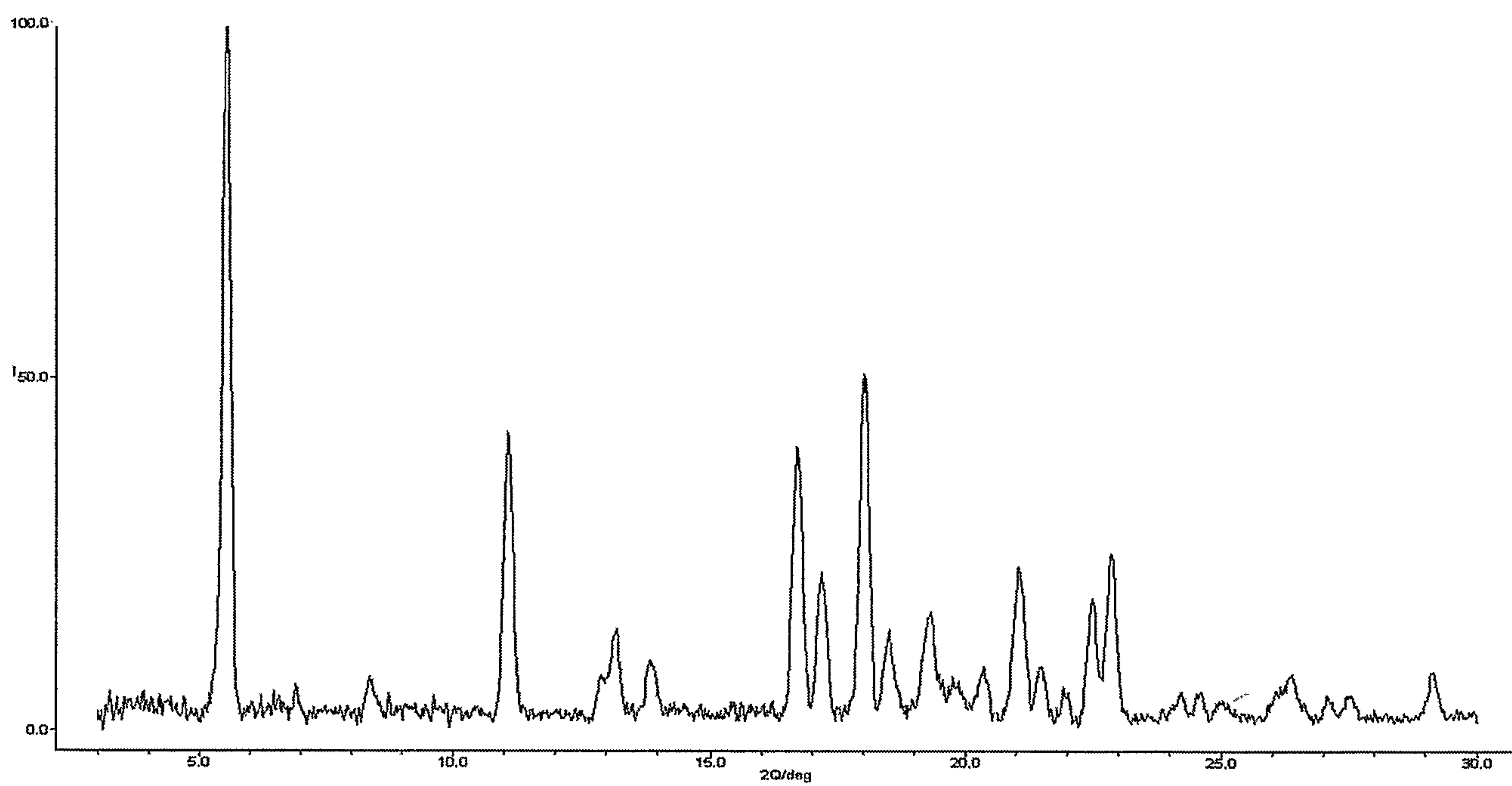


Figure 16

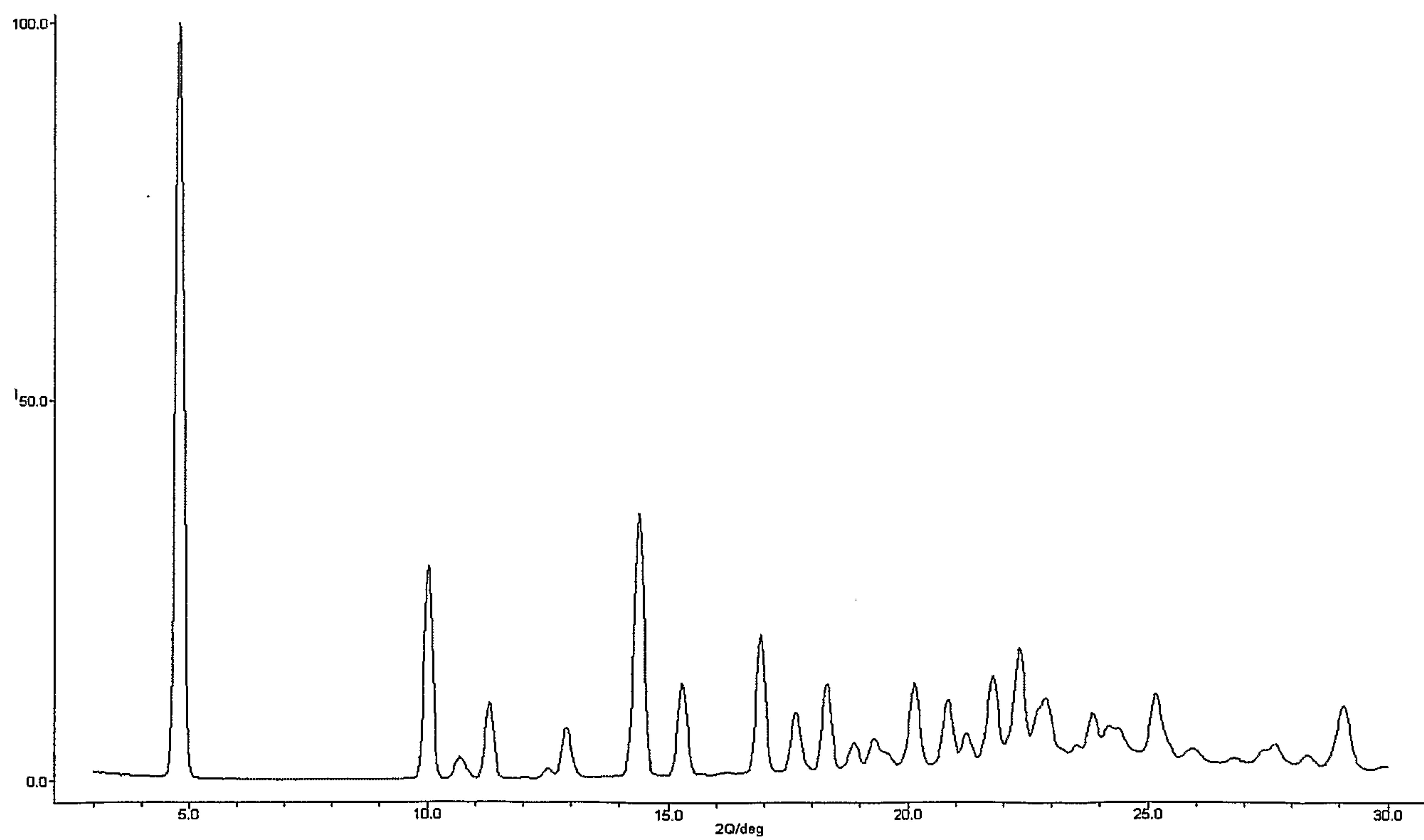


Figure 17

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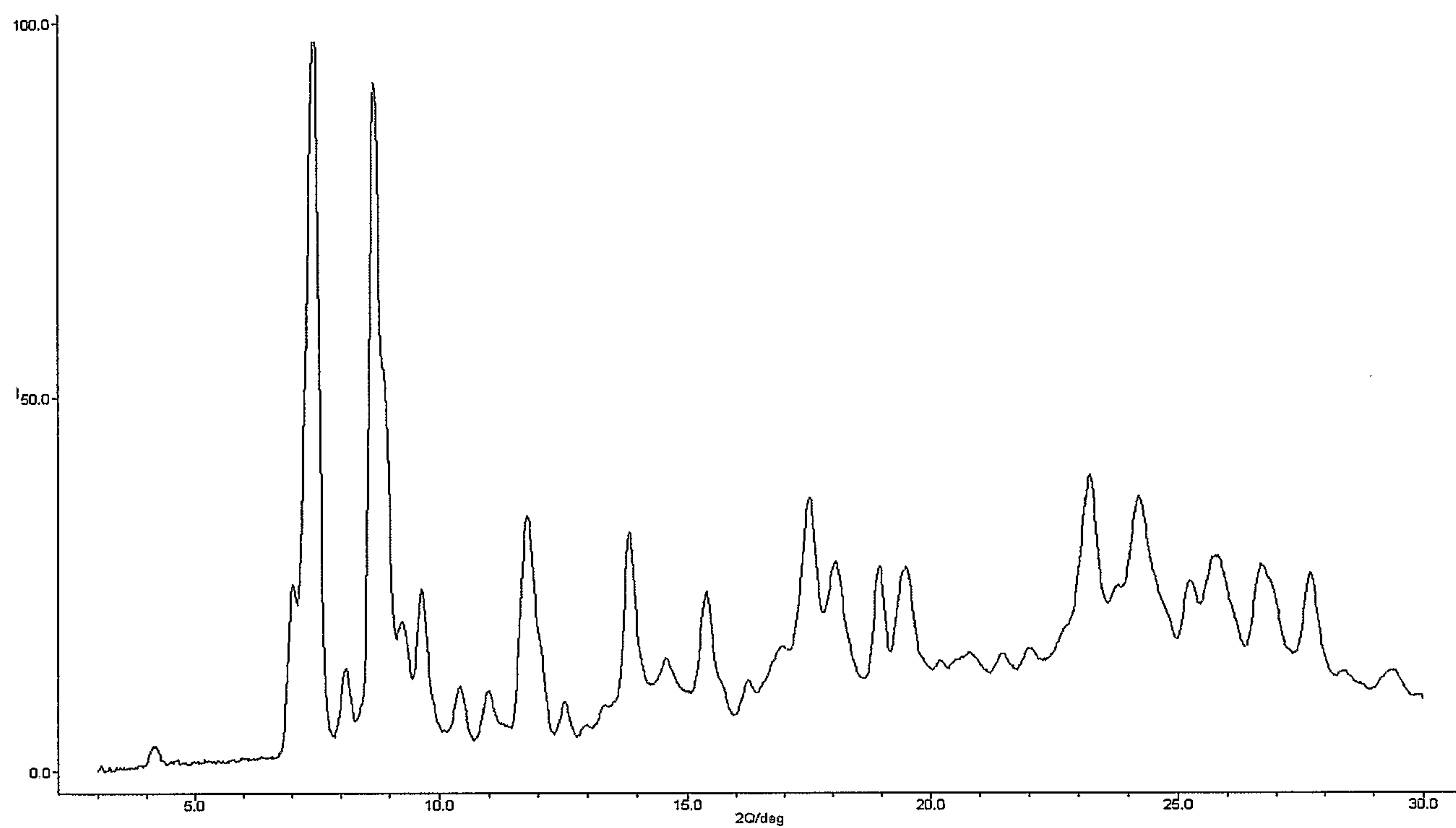


Figure 18

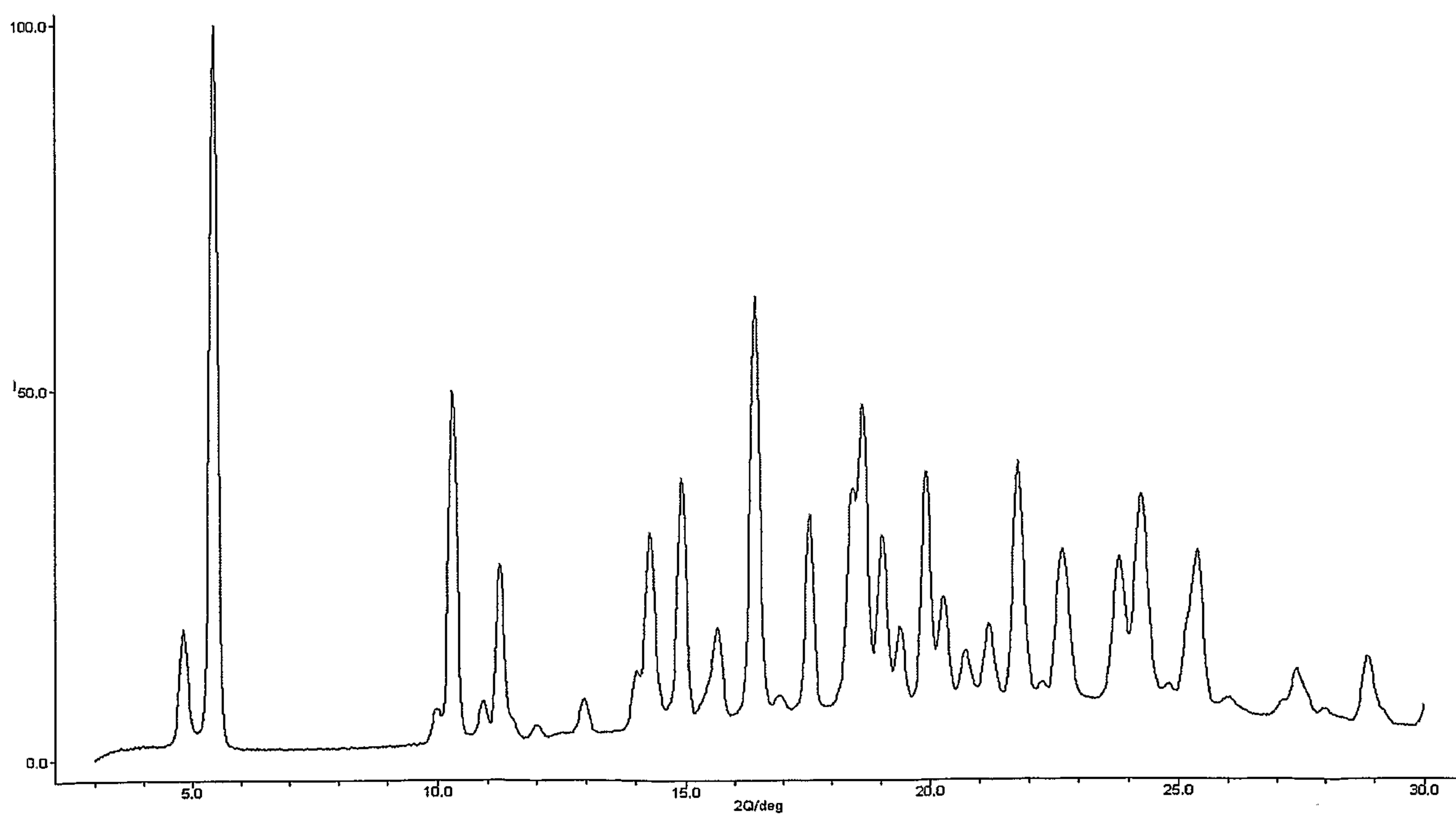


Figure 19

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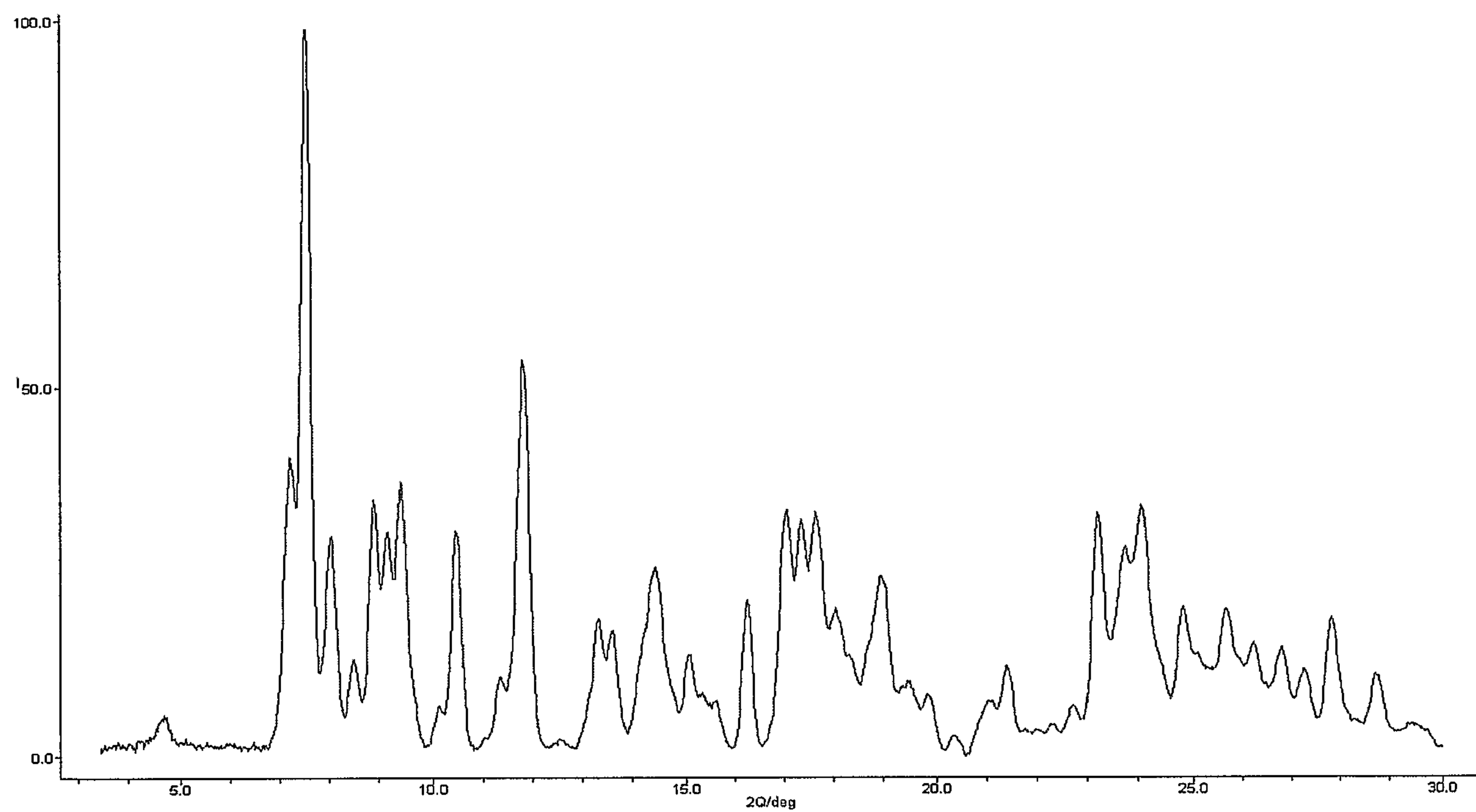


Figure 20

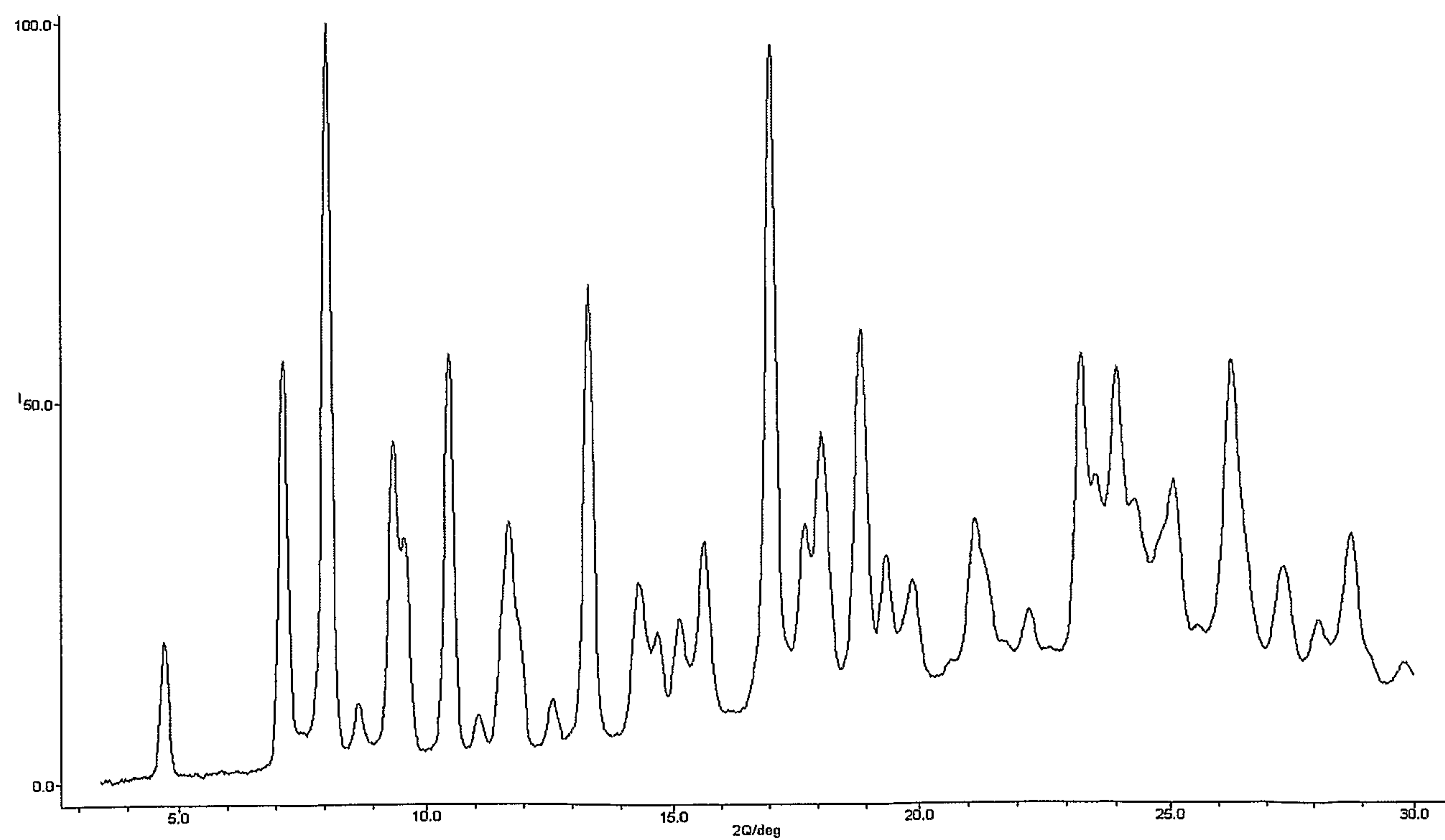


Figure 21

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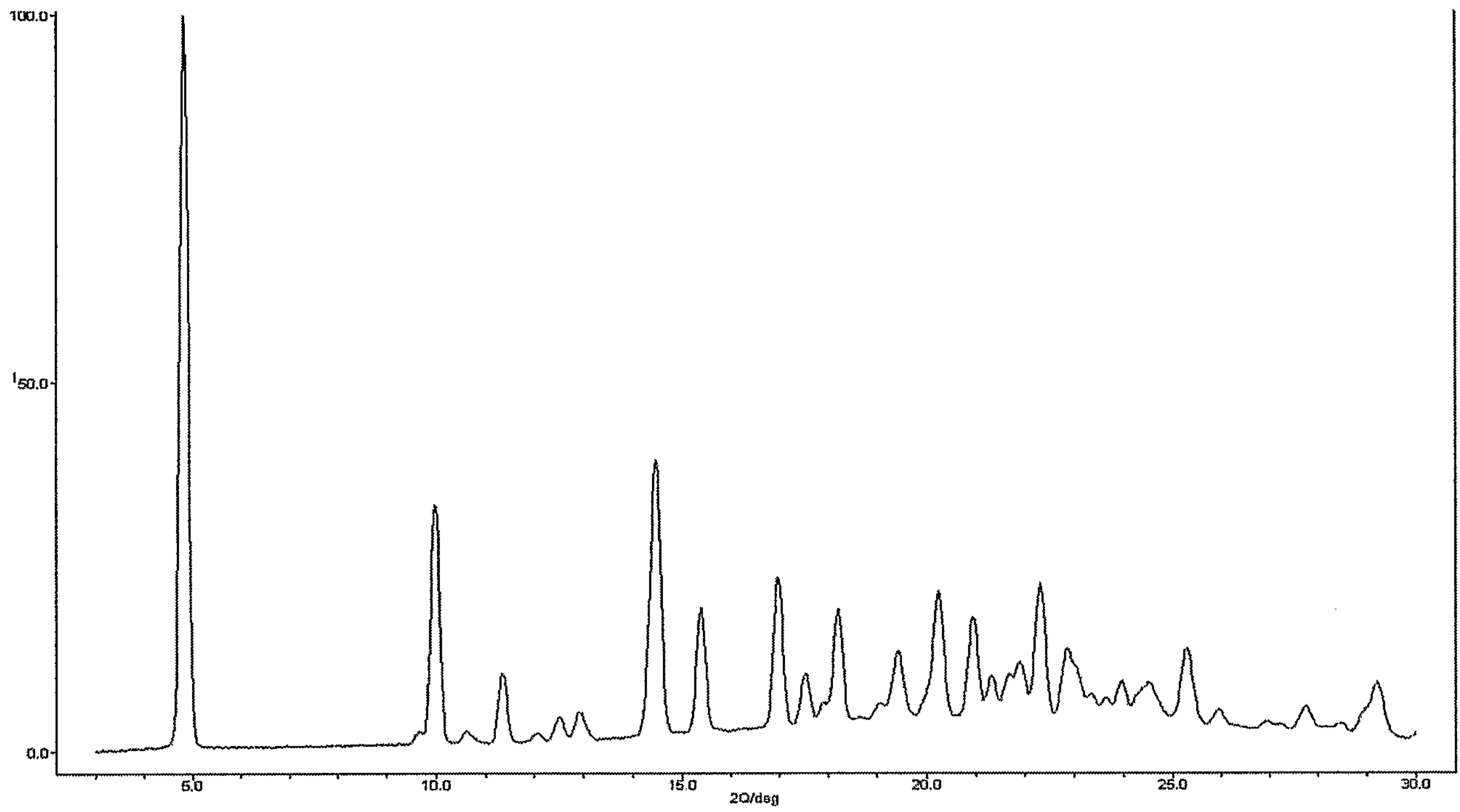
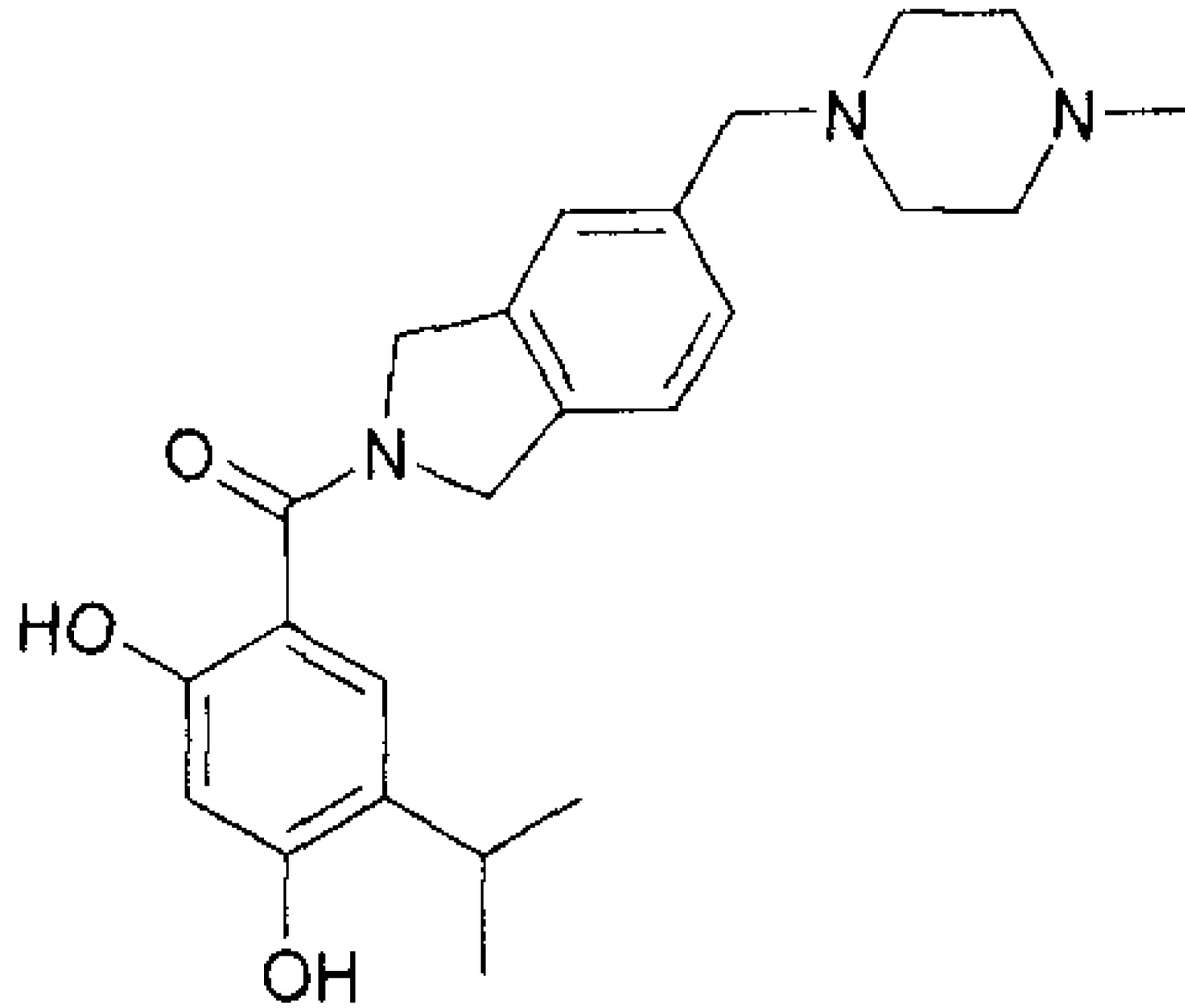


Figure 22



(1)