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(54) Title: FLUOROBETAINE COPOLYMER AND FIRE FIGHTING FOAM CONCENTRATES THEREFROM

(57) Abstract: It has been discovered by in present invention that, compared with a fluorosurfactant having only one betaine group per molecule, a fluorosurfactant having a plurality of betaine groups per molecule is superior in its ability to undergo electrostatic interaction with other components of a fire fighting composition thereby improving the performance of the fire fighting composition. Previously known fluorosurfactants comprise only one betaine group per molecule. In contrast, the present invention provides for a fluorobetaine copolymer containing a plurality of betaine groups.



WO 2008/115457 A2

TITLE OF INVENTION

FLUOROBETAINE COPOLYMER AND FIRE FIGHTING FOAM
CONCENTRATES THEREFROM

FIELD OF THE INVENTION

5 This invention relates to a fluorobetaine copolymer, the manufacture thereof, and the use thereof for use in a fire fighting foam concentrate.

BACKGROUND OF THE INVENTION

Fluorosurfactants which incorporate betaine groups are known for use as additives in fire extinguishing compositions. For example U.S. Pat. No. 6521730
10 discloses a fluorosurfactant having a betaine group represented by the formula $(N^+-CH_2-CO_2^-)$. The betaine group allows such fluorosurfactants to undergo electrostatic interactions with other components in a fire fighting composition therefore improving performance.

U.S. Pat. No. 6518345 which discloses a fire fighting composition
15 comprising a betaine containing fluorosurfactant, a polyethyleneimine, and a polybasic acid. The polybasic acid is necessary for the purpose of undergoing electrostatic interactions with the polyethyleneimine thereby improving the flame resistance and fuel resistance of the composition. Apparently, the fluorosurfactant alone could not sufficiently undergo electrostatic interaction with the
20 polyethyleneimine to improve the flame resistance and fuel resistance.

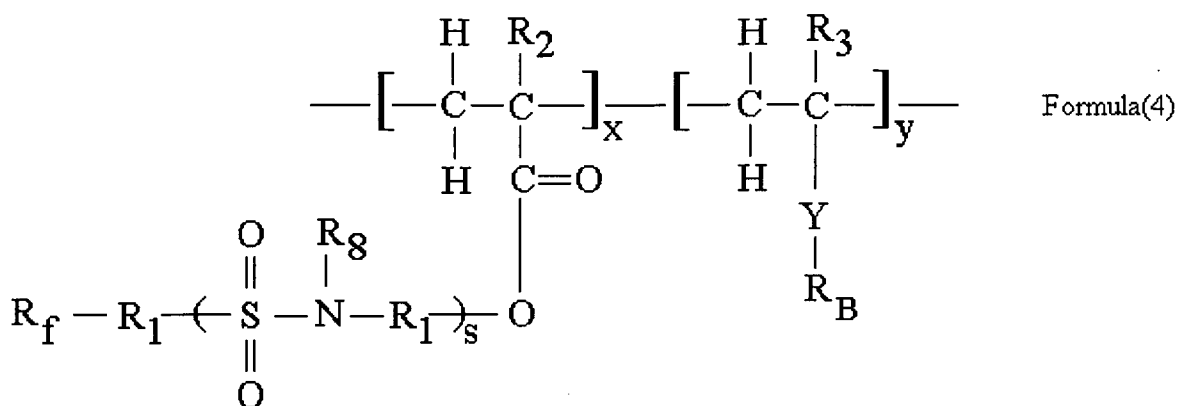
It would be desirable to discover a fluorosurfactant having an increased ability to undergo electrostatic interaction with other components of a fire fighting composition thereby improving the performance of the fire fighting composition.

BRIEF SUMMARY OF THE INVENTION

25 It has been discovered in the present invention that, compared with a fluorosurfactant having only one betaine group per molecule, a fluorosurfactant having a plurality of betaine groups per molecule is superior in its ability to undergo electrostatic interaction with other components of a fire fighting foam concentrate thereby improving the performance of the fire fighting foam
30 concentrate. Previously known fluorosurfactants, such as those disclosed in U.S. Pat. No. 6521730 and 6518345, comprise only one betaine group per molecule.

In contrast, the present invention provides for a fluorobetaine copolymer comprising a plurality of betaine groups denoted as R_B in the formula below.

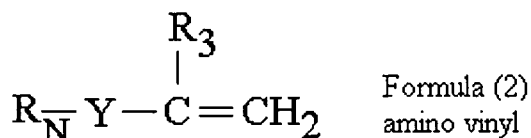
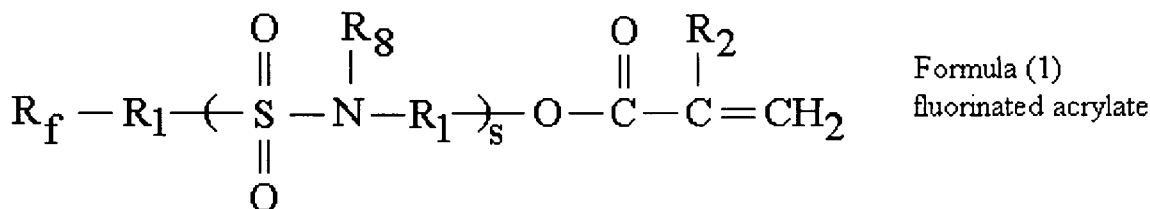
Fluorobetaine copolymers of the present invention comprise units represented by the following formula:



wherein R_f represents a perfluorinated alkyl group, optionally interrupted by oxygen, having at least 3 carbon atoms and preferably 6 or fewer carbon atoms; each R_1 is independently chosen from a C_1 - C_{20} hydrocarbylene, preferably C_1 - C_5 , more preferably linear, and even more preferably ethyl; R_8 is chosen from a C_1 - C_{20} hydrocarbyl, preferably C_1 - C_5 , more preferably linear, and even more preferably ethyl; R_2 is chosen from hydrogen, fluorine, or a C_1 - C_4 alkyl and is preferably hydrogen or methyl; s is 0 or 1, preferably 1; Y is benzyl or carbonyl, preferably carbonyl; R_B is a group with the formula $C_a N_b O_c S_f H_{2(a-e)+b+1-2d}$ having a structure which incorporates at least one betaine group of the formula ($N^+ - C_n H_{2n} - CO_2^-$) or ($N^+ - C_n H_{2n} - SO_3^-$) wherein a is an integer from 4-25, b is an integer from 1 to 4, c is an integer from 2 to 15, d integer from 0 to 3, e is an integer from 0 to 4, f is an integer from 0 to 4, the sum of e and f is from 1 to 4; n is independently chosen from an integer from 1 to 5; and R_3 is chosen from hydrogen or a C_1 - C_4 hydrocarbyl and is preferably hydrogen or methyl; and x is an integer from 1 to 10, y is an integer from 4 to 200, and the ratio of $y:x$ is at least 1:1, preferably at least 1:35, more preferably at least 1:45, and most preferably at least 1:60.

The fluorobetaine copolymer of the present invention can be made in two steps, the first step being the polymerization of monomers comprising at least one

fluorinated acrylate monomer and at least one amino vinyl monomer represented respectively by Formula (1) and Formula (2) as follows:



- 5 wherein R_f represents a perfluorinated alkyl group, optionally interrupted by oxygen, having at least 3 carbon atoms and preferably having 6 or fewer carbon atoms; each R_1 is independently chosen from a C_1 - C_{20} hydrocarbylene, preferably C_1 - C_5 , more preferably linear, and even more preferably ethyl; R_8 is chosen from a C_1 - C_{20} hydrocarbyl or hydrogen, preferably C_1 - C_5 , more preferably linear, and
- 10 even more preferably hydrogen or methyl; R_2 is chosen from hydrogen, fluorine, or a C_1 - C_4 alkyl and is preferably hydrogen or methyl; s is 0 or 1, preferably 1; R_N is a group incorporating at least one tertiary amine said group having the formula $\text{C}_a\text{N}_b\text{O}_c\text{H}_{2a+b+1-2d}$ where a is an integer from 3 to 20, preferably 3 to 15, b is a positive integer from 1 to 4, and c is an integer from 0 to 4, and d is an integer
- 15 from 0 to 4; and Y is carbonyl or benzyl, preferably carbonyl; and R_3 is chosen from hydrogen or a C_1 - C_4 hydrocarbyl and is preferably hydrogen or methyl.

- The second step of making the fluorobetaine copolymer is betaine conversion wherein a polymer obtained in the first step is reacted with a compound which converts at least a portion of the tertiary amines contained
- 20 therein. Betaine conversion can be accomplished by reacting the polymer obtained in the first step with at least one betaine precursor such as an aliphatic halocarboxylic acid having the formula $\text{X}-\text{A}-\text{COOH}$ wherein X is a halogen, preferably chlorine; A is a C_1 to C_4 alkylene, preferably C_1 to C_2 . Another example of a betaine precursor is an aliphatic halosulfonic acid having the
- 25 formula $\text{X}-\text{A}-\text{SO}_3\text{H}$ wherein X is a halogen, preferably chlorine; A is a C_1 to C_4 alkylene, preferably C_1 to C_2 . The aliphatic halocarboxylic acid and/or aliphatic halosulfonic acid is preferably used in the form of its salt, such as an

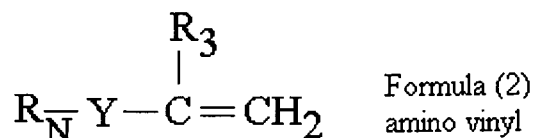
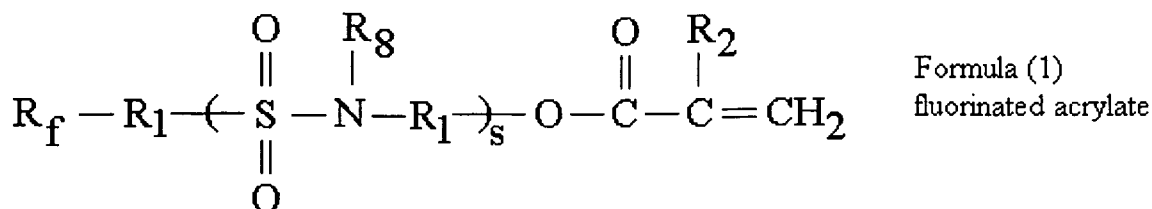
organic salt or alkaline earth metal salt, and more preferably its alkali metal salt. A betaine precursor particularly suitable for the invention is sodium chloroacetate. Betaine conversion may also be accomplished with a lactone, preferably a β -lactone such as β -propiolactone.

5 The fluorobetaine copolymers according to the invention lower the surface tension of aqueous solutions and are useful foaming agents and therefore are useful as additives in multipurpose fire fighting foam concentrate, in particular, fire fighting foam concentrates intended for combating hydrocarbon fires, such as fires in which the hydrocarbons are petrols, oils, diesel oil, fuel oil, heptane,
10 hexane or cyclohexane, or polar liquid fires, such as fires in which the polar liquids are alcohols (for example, methanol, ethanol and isopropanol), ketones (for example, dimethyl ketone and methyl isobutyl ketone), esters (for example, n-butyl acetate) and ethers (for example, methyl tert-butyl ether). The fluorobetaine copolymers are also useful as additives in multipurpose fire fighting foam
15 concentrate intended for combating Class A fires which are fires fueled by burning materials which leave an ash residue such as paper, wood, cloth, rubber, and certain plastics.

 The fire fighting foam concentrates of the present invention preferably include a polyethyleneimine. The fire fighting foam concentrates of the present
20 invention preferably include a hydrocarbon surfactant. The fire fighting foam concentrates of the present invention preferably include a monobasic acid, polybasic acid or mixtures thereof. Because of the betaine fluorocopolymer, the fire fighting foam concentrates are capable of excluding any polybasic acid.

25 **DETAILED DESCRIPTION OF THE INVENTION**

 The fluorobetaine copolymer of the present invention can be made in two steps, the first step being the polymerization of monomers comprising at least one fluorinated acrylate monomer and at least one amino vinyl monomer represented respectively by Formula (1) and Formula (2) as follows:

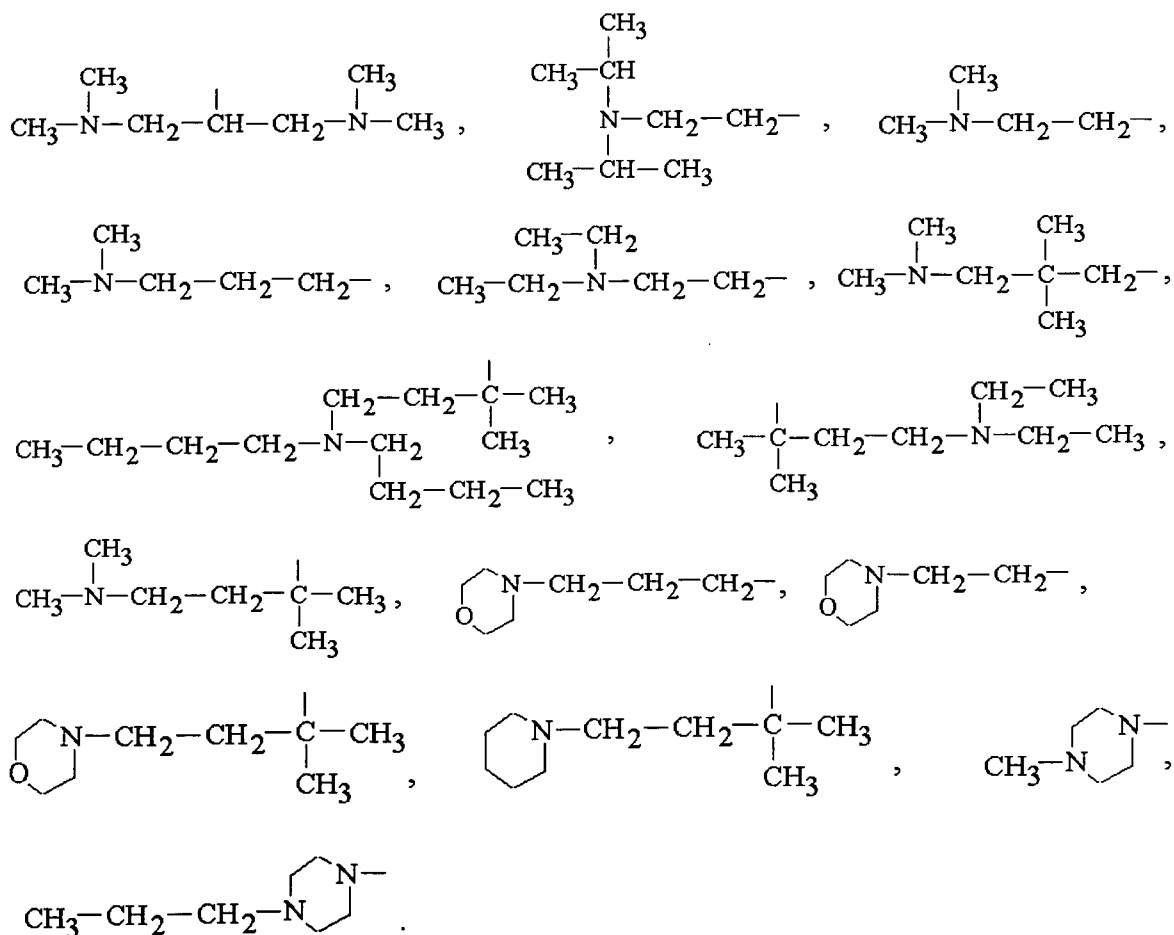


wherein R_f represents a perfluorinated alkyl group, optionally interrupted by oxygen, having at least 3 carbon atoms and preferably having 6 or fewer carbon atoms; each R_1 is independently chosen from a C_1 - C_{20} hydrocarbylene, preferably C_1 - C_5 , more preferably linear, and even more preferably ethyl; R_8 is chosen from a C_1 - C_{20} hydrocarbyl or hydrogen, preferably C_1 - C_5 , more preferably linear, and even more preferably hydrogen or methyl; R_2 is chosen from hydrogen, fluorine, or a C_1 - C_4 alkyl and is preferably hydrogen or methyl; s is 0 or 1, preferably 1; R_N is a group incorporating at least one tertiary amine said group having the formula $C_a N_b O_c H_{2a+b+1-2d}$ where a is an integer from 3 to 20, preferably 3 to 15, b is a positive integer from 1 to 4, and c is an integer from 0 to 4, and d is an integer from 0 to 4; and Y is carbonyl or benzyl, preferably carbonyl; and R_3 is chosen from hydrogen or a C_1 - C_4 hydrocarbyl and is preferably hydrogen or methyl. The term “hydrocarbyl” as used throughout is intended to mean a univalent group formed by removing a hydrogen atom from a hydrocarbon. The term “hydrocarbylene” as used throughout is intended to mean a divalent group formed by removing two hydrogens from a hydrocarbon, the free valencies of which are not engaged in a double bond.

Examples of suitable fluorinated acrylate monomers of Formula(1) wherein $s=0$ are disclosed in U.S. Pat. Nos. 4174851, 2642416, 3384627, 3392046, 3282905, 3532659, 3102103, all of which are hereby incorporated by reference to the extent permitted by applicable law. Specific examples of suitable fluorinated acrylate monomers of Formula(1) wherein $s=0$ include perfluorobutylethyl acrylate, perfluorohexylethyl acrylate, perfluorobutylethyl methacrylate, and perfluorohexylethyl methacrylate. Examples of suitable fluorinated acrylate monomers of Formula(1) wherein $s=1$ are disclosed in U.S.

- Pat. No. 5439998 which is hereby incorporated by reference to the extent permitted by applicable law. Specific examples of suitable fluorinated acrylate monomers of Formula(1) wherein s=1 include 2-
- [methyl[(3,3,4,4,5,5,6,6,7,7,8,8,8-tridecafluorooctyl)sulfonyl]amino]ethyl
- 5 acrylate; 2-[methyl[(3,3,4,4,5,5,6,6,6-nonafluorohexyl)sulfonyl]amino]ethyl
- acrylate; 2-[methyl[(2,2,3,3,4,4,5,5,6,6,7,7,7-
- tridecafluoroheptyl)sulfonyl]amino]ethyl acrylate; 2-
- [methyl[(3,3,4,4,5,5,6,6,7,7,8,8,8-tridecafluorooctyl)sulfonyl]amino]ethyl
- methacrylate; [methyl[(3,3,4,4,5,5,6,6,6-nonafluorohexyl)sulfonyl]amino]ethyl
- 10 methacrylate; and 2-[methyl[(2,2,3,3,4,4,5,5,6,6,7,7,7-
- tridecafluoroheptyl)sulfonyl]amino]ethyl methacrylate.

Examples Formula (2) depicting suitable R_N groups include:



- 15 Specific examples of amino vinyl monomers of Formula (2) include 2-
- (dimethylamino)-1-[(dimethylamino)methyl]ethyl acrylate, 3-(dimethylamino)-
- 2,2-dimethylpropyl acrylate, 2-(4-morpholinyl)ethyl acrylate, 2-(4-

morpholinyl)ethyl methacrylate, 2-(1-piperidinyl)ethyl methacrylate, 2-(1-piperidinyl)ethyl acrylate, 3-(dimethylamino)propyl acrylate, 2-[bis(1-methylethyl)amino]ethyl methacrylate, 2-(dimethylamino)ethyl acrylate, 2-(diethylamino)ethyl acrylate, 2-(diethylamino)ethyl methacrylate, N-[3-(dibutylamino)-1,1-dimethylpropyl]-2-methyl- methacrylamide, N-[3-(diethylamino)-1,1-dimethylpropyl]-2-methyl- methacrylamide, N-Acryloyl-N'-propylpiperazine, N-[1,1-dimethyl-3-(4-morpholinyl)propyl]-acrylamide, N-[1,1-dimethyl-3-(1-piperidinyl)propyl]-acrylamide, N-[3-(diethylamino)-1,1-dimethylpropyl]-acrylamide, 1-methyl-4-(2-methyl-1-oxo-2-propenyl)-piperazine, N-[3-(4-morpholinyl)propyl]- acrylamide, N-[3-(dimethylamino)-1,1-dimethylpropyl]-2-methyl- methacrylamide, N-[3-(dimethylamino)-1,1-dimethylpropyl]- acrylamide, N-[2-(diethylamino)ethyl]-2-methyl-methacrylamide, N-[2-(diethylamino)ethyl]-2-methyl- methacrylamide, N-[3-(dimethylamino)propyl]-2-methyl- methacrylamide, N-[3-(dimethylamino)propyl]- acrylamide, N-[2-(diethylamino)ethyl]- acrylamide, 4-Ethenyl-N,N-dimethylbenzylamine, and N-(p-Vinylbenzyl)dibutylamine.

The polymerization during the first step is free radical polymerization conducted in an organic solvent and enabled with a free radical initiator and a chain transfer agent for controlling molecular weight. Examples of free radical initiators include: azo compounds, such as azobisisobutyronitrile and azo-2-cyanovaleric acid; hydroperoxides, such as cumene, t-butyl and t-amyl hydroperoxide; dialkyl peroxides, such as di-t-butyl and dicumylperoxide; peroxyesters, such as t-butylperbenzoate and di-t-butylperoxy phthalate; and diacylperoxides, such as benzoyl peroxide and lauroyl peroxide. Examples of chain transfer agents include mercaptoethanol, mercaptoacetic acid, stearylmercaptane, tert-dodecylmercaptane, trichloromethane, diethyl phosphate, methanol and the like, and mixtures thereof. Examples of suitable organic solvents include: alcohols, such as 2-methylpropan-2-ol, isopropanol, 2-methoxypropan-2-ol; and ketones, such as acetone, methylisobutyl ketone, and methylethyl ketone; and pyrrolidones, such as n-methyl-2-pyrrolidone and the like, and mixtures thereof.

Free radical polymerization of monomers comprising a fluorinated acrylate monomer of Formula (1) and an amino vinyl monomer of Formula (2)

can be conducted by dissolving the monomers, free radical initiator, and chain transfer agent in an organic solvent. The solution can be heated and maintained at about 40 to 100° C, more preferably about 55 to 85° C, and allowed to react under inert conditions for a period of time to obtain at least 95 percent yield of polymer.

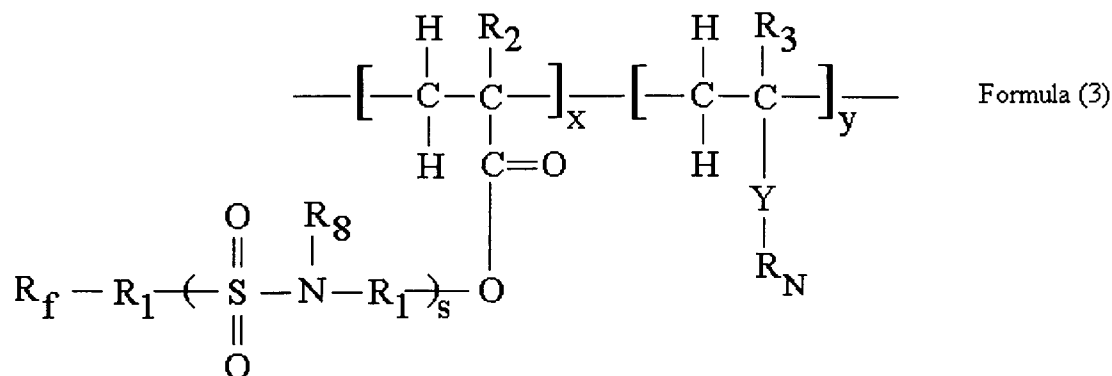
- 5 Polymer yield may be determined by measuring the amount of residual monomer by gas chromatography.

The concentration of monomers in the reaction solvent is preferably from 30 to 70 weight percent. The weight ratio of amino vinyl monomer to fluorinated acrylate monomer is preferably at least 1:1, more preferably at least 1:35, even
10 more preferably at least 1:45, and most preferably at least 1:60. Initiator is preferably added in an amount of 0.1 to 2 molar percentage of total monomers. Chain transfer agent(s) can be added in an amount to yield a polymer with a desirable targeted molecular weight which can be determined by summing the weight of monomers in grams and dividing the sum by the total moles of chain
15 transfer agent(s) used, and then adding to this quotient the weighted average of the molecular weight of the chain transfer agent(s) used. Preferable targeted molecular weights for fire extinguishing applications are no more than 40,000 $\text{g}\cdot\text{mol}^{-1}$, more preferably no more than 20,000 $\text{g}\cdot\text{mol}^{-1}$, and most preferably no more than 10,000 $\text{g}\cdot\text{mol}^{-1}$.

- 20 In the invention, free radical polymerization may also be conducted by “controlled free radical polymerization,” examples of which include: Atom Transfer Radical Polymerization (ATRP), Reversible Addition Fragmentation chain Transfer polymerization (RAFT), Iodine Transfer Polymerization (ITP), Reversible Iodine Transfer Polymerization (RITP), Macromolecular Design via
25 the Interchange of Xanthates (MADIX), and Nitroxide Mediated Polymerization (NMP). Of these mechanisms, RAFT and ITP are preferred for use in the invention.

The RAFT mechanism employs a dithioester group containing compounds as free radical initiators and is described in U.S. Pat No. 6,642,318; hereby
30 incorporated by reference to the extent permitted by applicable law. The ITP mechanism employs an iodo containing chain transfer agent and is described in U.S. Pat. Nos. 4,158,678; 5,231,154; both of which are hereby incorporated by reference to the extent permitted by applicable law.

The polymer obtained after the free radical polymerization of the first step comprises units represented by Formula (3) as follows:



wherein R_f represents a perfluorinated alkyl group, optionally interrupted by oxygen, having at least 3 carbon atoms and preferably having 6 or fewer carbon atoms; each R_1 is independently chosen from a C_1 - C_{20} hydrocarbylene, preferably C_1 - C_5 , more preferably linear, and even more preferably ethyl; R_8 is chosen from a C_1 - C_{20} hydrocarbyl or hydrogen, preferably C_1 - C_5 , more preferably linear, and even more preferably hydrogen or methyl; R_2 is chosen from hydrogen, fluorine, or a C_1 - C_4 alkyl and is preferably hydrogen or methyl; s is 0 or 1, preferably 1; R_N is a group incorporating at least one tertiary amine said group having the formula $C_a N_b O_c H_{2a+b+1-2d}$ where a is an integer from 3 to 20, preferably 3 to 15, b is a positive integer from 1 to 4, and c is an integer from 0 to 4, and d is an integer from 0 to 4; and Y is carbonyl or benzyl, preferably carbonyl; and R_3 is chosen from hydrogen or a C_1 - C_4 hydrocarbyl and is preferably hydrogen or methyl; and x is an integer from 1 to 10, y is an integer from 4 to 200, and the ratio of $y:x$ is at least 1:1, preferably at least 1:35, more preferably at least 1:45, and most preferably at least 1:60.

The second step of making the fluorobetaine copolymer is betaine conversion wherein a polymer of Formula (3) is reacted with a compound which converts at least a portion of the tertiary amines in the pendant groups identified as R_N . Betaine conversion can be accomplished by compounds well known in the art for reacting with tertiary amine groups to form betaine groups.

Betaine conversion can be accomplished by reacting the polymer obtained in the first step, i.e., Formula (3), with at least one betaine precursor such as an aliphatic halocarboxylic acid having the formula $X\text{---}A\text{---}COOH$ wherein X is a

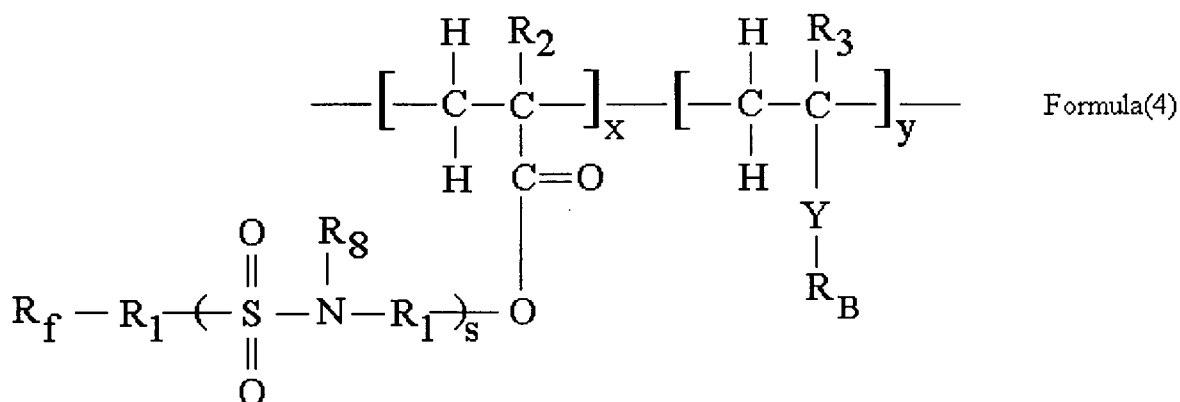
halogen, preferably chlorine; A is a C₁ to C₄ alkylene, preferably C₁ to C₂.

Another example of a betaine precursor is an aliphatic halosulfonic acid having the formula X—A—SO₃H wherein X is a halogen, preferably chlorine; A is a C₁ to C₄ alkylene, preferably C₁ to C₂. The aliphatic halocarboxylic acid and/or

- 5 aliphatic halosulfonic acid is preferably used in the form of its salt, such as an organic salt or alkaline earth metal salt, and more preferably its alkali metal salt. The term "alkylene" as used throughout is intended to mean an acyclic carbon or a saturated acyclic carbon chain represented by the formula C_nH_{2n}. An aqueous solution of the betaine precursor can be mixed with a polymer of Formula (3)
- 10 dissolved in an organic solution. This organic solution conveniently can be the medium in which the polymer was made. The mixture can be heated to about 40 to 100° C, more preferably about 60 to 90° C, for a period of time preferable to convert significantly all of the tertiary amines into betaines, typically over about 1 hour. The organic solvent can then be removed by distillation.

- 15 Betaine conversion may also be accomplished with a lactone, preferably a β-lactone such as β-propiolactone.

Fluorobetaine copolymers of the present invention comprise units represented by Formula (4) as follows:



- 20 wherein R_f represents a perfluorinated alkyl group, optionally interrupted by oxygen, having at least 3 carbon atoms and preferably having 6 or fewer carbon atoms; each R₁ is independently chosen from a C₁-C₂₀ hydrocarbylene, preferably C₁-C₅, more preferably linear, and even more preferably ethyl; R₈ is chosen from a C₁-C₂₀ hydrocarbyl or hydrogen, preferably C₁-C₅, more preferably linear, and
- 25 even more preferably hydrogen or methyl; R₂ is chosen from hydrogen, fluorine,

or a C₁-C₄ alkyl and is preferably hydrogen or methyl; s is 0 or 1, preferably 1; Y is benzyl or carbonyl, preferably carbonyl; R_B is a group with the formula C_aN_bO_cS_fH_{2(a-e)+b+1-2d} having a structure which incorporates at least one betaine group of the formula (N⁺-C_nH_{2n}-CO₂⁻) or (N⁺-C_nH_{2n}-SO₃⁻) wherein a is an integer from 4-25, b is an integer from 1 to 4, c is an integer from 2 to 15, d integer from 0 to 3, e is an integer from 0 to 4, f is an integer from 0 to 4, the sum of e and f is from 1 to 4; n is independently chosen from an integer from 1 to 5; and R₃ is chosen from hydrogen or a C₁-C₄ hydrocarbyl and is preferably hydrogen or methyl; and x is an integer from 1 to 10, y is an integer from 4 to 200, and the ratio of y:x is at least 1:1, preferably at least 1:35, more preferably at least 1:45, and most preferably at least 1:60.

The fluorobetaine copolymers according to the invention lower the surface tension of aqueous solutions and are useful foaming agents and therefore are useful as additives in multipurpose fire fighting foam concentrates, in particular, fire fighting foam concentrates intended for combating hydrocarbon fires, such as fires in which the hydrocarbons are petrols, oils, diesel oil, fuel oil, heptane, hexane or cyclohexane, or polar liquid fires, such as fires in which the polar liquids are alcohols (for example, methanol, ethanol and isopropanol), ketones (for example, dimethyl ketone and methyl isobutyl ketone), esters (for example, n-butyl acetate) and ethers (for example, methyl tert-butyl ether). The fluorobetaine copolymers are also useful as additives in multipurpose fire fighting foam concentrates or compositions intended for combating Class A fires which are fires fueled by burning materials which leave an ash residue such as paper, wood, cloth, rubber, and certain plastics.

Fire fighting foam concentrates are liquid compositions intended for combating combustible liquid fires, particularly those caused by hydrocarbons and/or polar liquids). Typically at the time of use, the foam concentrates are diluted in municipal water or seawater, generally at a concentration by weight of 3% (that is to say, 3 parts by weight of foam concentrates per 97 parts by weight of water) or 6% (6 parts by weight of foam concentrate per 94 parts by weight of water) but also, and more rarely 1% (1 part by weight of foam concentrate per 99 parts by weight of water). The foaming solution can be obtained by diluting the foam concentrate with water. This foaming solution passes through a fire hose

nozzle where mechanical agitation takes place with incorporation of air, which generates an extinguishing foam used to combat combustible liquid fires.

When the fluorobetaine copolymers according to the invention are incorporated in the foam concentrates, they improve the stability of the extinguishing foams on a polar liquid and thus their extinguishing performance on fires of this type. Their content in the foam concentrates can generally range from 0.1 to 10 percent by weight and preferably from 0.2 to 5 percent by weight.

The foam concentrates in which the fluorobetaine copolymers can be incorporated are of two types, depending upon the origin of their foaming base. Synthetic foam concentrates, the foaming base of which is composed of at least one hydrocarbonaceous surface-active agent, and protein foam concentrates, the foaming base of which is composed of an animal protein hydrolysate, are distinguished. These two types of foam concentrates can comprise, according to their destination, one or more fluorosurfactants, one or more foam-stabilizing cosolvents, a hydrophilic polymer of high molecular weight of the polysaccharide type with a thixotropic and alcohol-repellent nature, an antifreezing agent, a corrosion-inhibiting agent, a preservative, a pH stabilizer or inorganic salts in which the cation is divalent, such as, for example, the magnesium ion or the calcium ion

In addition to the fluorobetaine copolymer of Formula (4), the fire fighting foam concentrates of the present invention preferably include a polyethyleneimine. The term "polyethyleneimine" as used herein includes polyethyleneimine and derivatives thereof. Examples of useful polyethyleneimines are disclosed in U.S. Pat. No. 6518345 hereby incorporated by reference to the extent permitted by applicable law. Specific examples of the polyethyleneimine useful in the fire extinguishing composition identified by their trade names include: LUPASOL PS with an average molecular weight of about $750,000 \text{ g}\cdot\text{mol}^{-1}$, solids content of about 31.5-34.5 weight percent, commercially available from BASF AG; LUPASOL HF with an average molecular weight of about $25,000 \text{ g}\cdot\text{mol}^{-1}$, solids content of about 56 weight percent, commercially available from BASF AG; and EPOMIN 1050 with an average molecular weight of about $70,000 \text{ g}\cdot\text{mol}^{-1}$, solids content of about 50 weight percent, commercially available from Nippon Sokubai. The amount of polyethyleneimine in the fire

fighting foam concentrates of the present invention can generally range from 0.1 to 15 percent by weight and preferably from 0.2 to 12 percent by weight.

In known fire fighting compositions, such as U.S. Pat. No. 6518345, a polybasic acid compound is utilized to undergo electrostatic interaction with a polyethyleneimine for the purpose of improving flame resistance and fuel resistance when used as fire fighting foam concentrates. Unlike these known fire fighting compositions, the fire fighting foam concentrates of the present invention do not require the use of a polybasic acid compound because the multiple betaine groups in the fluorobetaine copolymers of Formula (4) undergo an electrostatic interaction with polyethyleneimine and therefore can completely replace any polybasic acid while providing flame resistance and fuel resistance when used as fire fighting foam concentrates.

In addition to the fluorobetaine copolymer of Formula(4), the fire fighting foam concentrates of the present invention preferably include at least one hydrocarbon surfactant which can be cationic, anionic, non-ionic, amphoteric, or mixtures thereof. Examples of hydrocarbon surfactant include cetyltrimethylammonium bromide, quaternized cocoamine ethoxylate, Triton X100 (trade name of Union Carbide), and dioctyl sodium sulfosuccinate. The amount of hydrocarbon surfactant in the fire fighting foam concentrates of the present invention can generally range from 0.1 to 20 percent by weight and preferably from 0.2 to 15 percent by weight.

EXAMPLES

The following describes the materials used in the examples below.

LUPASOL PS (BASF AG trade name) is a polyethyleneimine with an average molecular weight of 750,000 mol/g and a solids content of 31.5-34.5 weight percent. EPOMIN 1050 (Nippon Sokubai trade name) is a polyethyleneimine with an average molecular weight of 70,000 mol/g and a solids content of 50 weight percent. AMPHOTENSID B4F (Zschimmer & Schwarz trade name) is a hydrocarbon surfactant with a solids content of 35 weight percent. The fluorinated acrylate monomer used was 2-[methyl[(3,3,4,4,5,5,6,6,7,7,8,8,8-tridecafluorooctyl sulfonyl]amino)ethyl acrylate. The comparative monomeric fluorosurfactant surfactant used is made in

accordance with U.S. Pat. No. 3721706 and comprises molecules each having a perfluorinated alkyl chain and one betaine group.

Extinguishing times were measured according to the following procedure. 150 mL of acetone was poured into a circular metal container with an internal diameter of 115mm. Moreover, an aqueous solution composed of the fire fighting foam concentrate diluted to 6 percent by weight in tap water was prepared. This solution was the foaming solution. A rotary stirrer composed of a motor and a metal rod, at the end of which were attached paddles produced a mechanical effect when the rod was being rotated; the rotational speed is adjustable from 0 to 2,800 rpm. The rod was introduced into the bottom of a cylindrical container equipped with an inlet orifice located at the bottom and with an outlet orifice located at the top. A metering pump transfers, via the inlet orifice, the aqueous solution to the bottom of the cylindrical container; foam was produced on contact with the rotating paddles, which foam was discharged as it was formed, via the outlet orifice. The throughput of the pump and the rotational speed of the rod were adjusted so that foam is continuously produced with a stationary foam throughput equal to about 40 g per minute. When the foam throughput was stabilized, the acetone was ignited. After the acetone burned for 90 seconds, the foam was poured into the metal container via a single point situated on the circumference. When the acetone was completely extinguished, the extinguishing time was recorded. The foam concentrates with the best performance on polar solvent were those for which the extinguishing time was as low as possible.

Re-ignition times were measured according to the following procedure. This parameter can be evaluated if the extinguishing time was less than 120 seconds. In this case, the foam was poured over the acetone even after the fire seat was extinguished. In all the cases described here, the foam was poured for 120s. 60s after the pouring of the foam was halted, the contents of re-ignition vessels (metal container with a diameter of 55mm and height of 40mm filled with acetone to a height of 20mm) was ignited. The re-ignition vessel was placed at the center of the metal container described above, the surface of fuel present in the said container being kept covered with foam. The time at the end of which the flames destroyed 25% of the surface initially covered the foam and spread in a

lasting fashion over the surface of the metal container was recorded. The greater this time, the better the ability of the foam to prevent the resurgence of the fire.

Example 1

5 A fluorobetaine copolymer of the invention was made according to the following procedure. To a double-jacketed 2.5-L reactor was added dimethylaminoethylmethacrylate (183.8 g, 1.17 mol), fluorinated acrylate monomer (81.8 weight percent purity in acetone, 110.0 g, 0.167 mol), t-butanol (422.9 g, 5.71 mol), and n-dodecyl mercaptan (19.9 g, 97.9 mmol) while stirring
10 (200 rpm). The temperature was raised to 70°C. The reactor was purged with a nitrogen flow for 25 minutes. Azobisisobutyronitrile (1.42 g, 8.64 mmol) was added to the solution. The temperature was maintained for 4 hours under nitrogen. Azobisisobutyronitrile (0.35 g, 2.13 mmol) was added to the solution and the temperature was maintained for 3 hours under nitrogen. A solution of
15 sodium chloroacetate (141.6 g, 1.22 mol) in water (1303.4 g) was added drop by drop over the course of 1 hour and 10 minutes. The temperature was maintained at 70°C for 1 hour. From this mixture, 500g of solvent was removed by distillation. The solid content of the solution was 26.3 weight percent including 4.16 weight percent of NaCl and the pH was 4.1. The targeted molecular weight
20 of the fluorobetaine copolymer was 3,700 g•mol⁻¹.

Example 2

 A fire fighting foam concentrate of the invention was made according to the following procedure. 33g of tap water, 18g of LUPASOL PS, 2.85g of
25 AMPHOTENSID B4F, 9.1g of the product of Example 1 was added at room temperature under vigorous stirring, the pH is adjusted to 7.5 by diglycolic acid, and the solution was adjusted up to 100g with tap water. A foaming solution was obtained by a dilution of this adjusted solution at 6 percent by weight in tap water to obtain a solution containing 144ppm of fluorine. The resulting solution,
30 subjected to the tests described above, exhibited the following characteristics: 1 minutes and 27 seconds of extinguishing time; and 6 minutes 10 seconds of re-ignition time.

Comparative Example A

A comparative fire fighting foam concentrate was made according to the following procedure. 33g of tap water, 18g of LUPASOL PS, 2.85 of AMPHOTENSID B4F, and 5.0g of comparative monomeric fluorosurfactant (solid content: 27 weight percent) were added at room temperature under vigorous stirring, the pH was adjusted to 7.5 by diglycolic acid, and the solution was adjusted up to 100g with tap water. A foaming solution was obtained by a dilution of this solution at 6 percent by weight in tap water in order to obtain a solution containing 225ppm of fluorine. The resulting solution, subjected to the tests described above, exhibited the following characteristics: greater than 4 minutes of extinguishing time.

Example 3

A fire fighting foam concentrate of the invention was made according to the following procedure. 33g of tap water, 12g of EPOMIN P-1050, 2.85 of AMPHOTENSID B4F, and 9.1g of the product of Example 1 was added at room temperature under vigorous stirring, the pH was adjusted to 7.5 by diglycolic acid, and the solution was adjusted up to 100g with tap water. A foaming solution was obtained by a dilution of this solution at 6 percent by weight in tap water in order to obtain a solution containing 144ppm of fluorine. The resulting solution, subjected to the tests described above, exhibited the following characteristics: 1 minute and 32 seconds of extinguishing time; and 5 minutes and 48 seconds of re-ignition time.

Comparative Example B

A comparative fire fighting foam concentrate was made according to the following procedure. 33g of water, 12g of EPOMIN P-1050, 2.85g of AMPHOTENSID B4F, and 1.96g of comparative monomeric fluorosurfactant (solid content: 27 weight percent) were added at room temperature under vigorous stirring, the pH was adjusted to 7.5 by diglycolic acid, and the solution was adjusted up to 100g with tap water. A foaming solution was obtained by a dilution of this solution at 6 percent by weight in tap water in order to obtain a solution containing 144ppm of fluorine. The resulting solution, subjected to the

tests described above, exhibited the following characteristics: greater than 4 minutes of extinguishing time.

Example 4

- 5 A fire fighting foam concentrate of the invention was made according to the following procedure. 33g of tap water, 20g of EPOMIN P-1050, 2.85g of AMPHOTENSID B4F, and 9.1g of the product of Example 1 were added at room temperature under vigorous stirring, the pH was adjusted to 7.5 by chloridric acid, and the solution was adjusted up to 100g with tap water. A foaming solution was
10 obtained by a dilution of this solution at 6 percent by weight in tap water in order to obtain a solution containing 144ppm of fluorine. The resulting solution, subjected to the tests described above, exhibited the following characteristics: 1 minute and 36 seconds of extinguishing time; and 7 minutes and 2 seconds of re-ignition time.
- 15 In the examples above: Example A is a comparative example of Example 2; and Example B is a comparative Example of 3. In these examples, an evaluation was preformed comparing the effect of fluorobetaine copolymers (Examples 2 and 3) versus the effect of comparative monomeric fluorosurfactant (Examples A and B) upon extinguishing times. The fluorobetaine copolymer and
20 comparative monomeric fluorosurfactant were analogous in that both have betaine groups. However the fluorobetaine copolymers have multiple betaine groups per molecule while the comparative monomeric surfactant only had a single betaine group per molecule. Thus the fluorobetaine copolymers had a greater ability to undergo electrostatic interaction with other components which resulted in
25 improved fire fighting performance. Evidence of this can be seen in Example 2 where the fluorobetaine copolymer yielded an extinguishing time of 2 minutes and 17 seconds. Under similar conditions in Example A, the comparative monomeric fluorosurfactant yielded an extinguishing time greater than 4 minutes. Further evidence can be seen in Example 3 where the fluorobetaine copolymer
30 yielded an extinguishing time of 1 minute and 32 seconds. Under similar conditions in Example 3A, the comparative monomeric fluorosurfactant yielded an extinguishing time greater than 4 minutes.

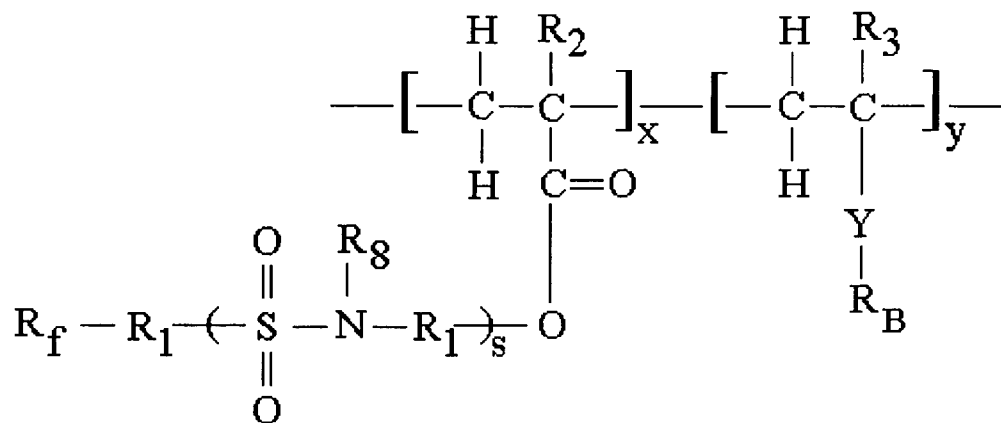
Although Example 3 and 4 both represent a fire fighting foam concentrate of the invention and both use a fluorobetaine copolymer of the invention, a polybasic acid compound was used in Example 4 while a monobasic acid compound was used in Example 3. As demonstrated by the similar extinguishing and re-ignition times of Examples 3 and 4, a polybasic acid was not required to impart fire fighting performance when using a fluorobetaine copolymer of the invention. In contrast, known fire fighting compositions, such as U.S. Pat. No. 6518345, require a polybasic acid to undergo electrostatic interaction with a polyethyleneimine for the purpose of improving fire fighting performance. The present invention does not require the use of a polybasic acid compound because the multiple betaine groups in the fluorobetaine copolymers of Formula(4) undergo an electrostatic interaction with polyethyleneimine and therefore can completely replace any polybasic acid while providing flame resistance and fuel resistance when used as fire fighting foam concentrates.

15

CLAIM(S)

What is claimed is:

1. A fluorocopolymer comprising units represented by the following formula:



- 5 wherein R_f represents a perfluorinated alkyl group, optionally interrupted by oxygen, having at least 3 carbon atoms; each R_1 is independently chosen from a C_1 - C_{20} hydrocarbylene; R_8 is chosen from a C_1 - C_{20} hydrocarbyl or hydrogen; R_2 is chosen from hydrogen, fluorine, or a C_1 - C_4 ; s is 0 or 1; Y is benzyl or carbonyl, preferably carbonyl; R_B is a group with the formula $C_a N_b O_c S_f H_{2(a-e)+b+1-2d}$ having a structure which incorporates at least one betaine group of the formula (N^+ - $C_n H_{2n} \text{---} CO_2^-$) or (N^+ - $C_n H_{2n} \text{---} SO_3^-$) wherein a is an integer from 4-25, b is an integer from 1 to 4, c is an integer from 2 to 15, d integer from 0 to 3, e is an integer from 0 to 4, f is an integer from 0 to 4, the sum of e and f is from 1 to 4; n is
- 10 independently chosen from an integer from 1 to 5; and R_3 is chosen from hydrogen or a C_1 - C_4 hydrocarbyl; and x is an integer from 1 to 10, y is an integer from 4 to 200, and the ratio of $y:x$ is at least 1:1.

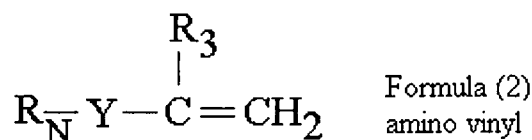
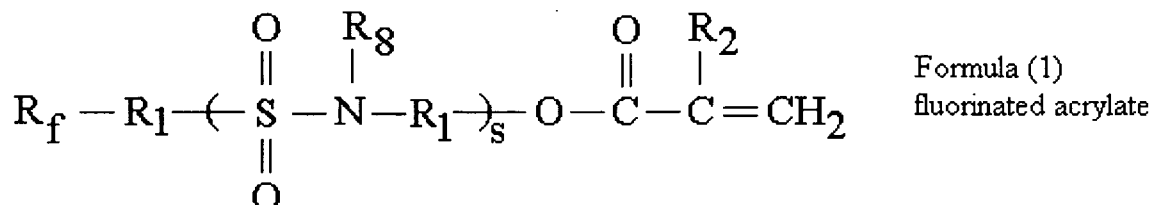
2. The fluorocopolymer of claim 1 wherein s is 1.

3. The fluoropolymer of any claims 1-2 wherein R_f has from 1 to 6 carbon atoms; each R_1 is independently a C_1 - C_5 hydrocarbylene; R_8 is chosen from C_1 - C_5 hydrocarbyl or hydrogen; R_2 is chosen from hydrogen or methyl; Y is carbonyl; R_3 is chosen from hydrogen or methyl; and the ratio of $y:x$ is at least 1:35.

4. A method for making the fluorocopolymer of any claims 1-3 comprising the steps of:

A) polymerizing of monomers comprising at least one fluorinated acrylate monomer and at least one amino vinyl monomer represented respectively by

5 Formula (1) and Formula (2) as follows:



wherein R_f represents a perfluorinated alkyl group, optionally interrupted by oxygen, having at least 3 carbon; each R_1 is independently chosen from a C_1 - C_{20}

10 hydrocarbylene, preferably C_1 - C_5 ; R_8 is chosen from a C_1 - C_{20} hydrocarbyl;

R_2 is chosen from hydrogen, fluorine, or a C_1 - C_4 alkyl; R_8 is chosen from a C_1 -

C_{20} hydrocarbyl or hydrogen; R_N is a group incorporating at least one tertiary amine said group having the formula $\text{C}_a\text{N}_b\text{O}_c\text{H}_{2a+b+1-2d}$ where a is an integer from 3 to 20, b is a positive integer from 1 to 4, and c is an integer from 0 to 4, and d is

15 an integer from 0 to 4; and Y is carbonyl or benzyl; and R_3 is chosen from hydrogen or a C_1 - C_4 hydrocarbyl; and

B) reacting the polymer obtained in step A) with at least one betaine forming compound chosen from:

- 20 i) an aliphatic halocarboxylic acid having the formula $\text{X}-\text{A}-\text{COOH}$ wherein X is a halogen, A is a C_1 to C_4 alkylene;
- ii) an aliphatic halosulfonic acid having the formula $\text{X}-\text{A}-\text{SO}_3\text{H}$ wherein X is a halogen; A is a C_1 to C_4 alkylene; or
- iii) a lactone.

25

5. The method of claim 4 wherein s is 1.

6. The method of any claims 4-5 wherein the aliphatic halocarboxylic acid or aliphatic halosulfonic acid is used in the form of its neutral alkaline earth metal salt.
- 5 7. A fire fighting foam concentrate comprising the fluorocopolymer of any claims 1-3.
8. The fire fighting foam concentrate of claim 7 further comprising a polyethyleneimine.
- 10 9. The fire fighting foam concentrate of any claims 7-8 further comprising a hydrocarbon surfactant.
10. The fire fighting foam concentrate of any claims 8-9 further comprising a monobasic acid, polybasic acid or mixtures thereof.
- 15 11. The use of any claims 1-3 as a fire fighting foam concentrate.