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(54) Title: METALLIFEROUS STORAGE MATERIAL FOR HYDROGEN AND METHOD FOR PRODUCING SAME

(57) **Abrégé/Abstract:**

The invention relates to a metalliferous material and a method for producing the same. The metalliferous material contains at least one metal oxide as catalyst means for the hydrogenation or dehydrogenation of said metalliferous material. According to the inventive method for producing such a metalliferous material, i.e. the actual storage material, the metalliferous material and/or the catalyst means are subjected to a mechanic milling process.





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(54) Title: METALLIFEROUS STORAGE MATERIAL FOR HYDROGEN AND METHOD FOR PRODUCING SAME (54) Bezeichnung: METALLHALTIGER WASSERSTOFFSPEICHERWERKSTOFF UND VERFAHREN ZU SEINER HERSTELLUNG (57) Abstract The invention relates to a metalliferous material and a method for producing the same. The metalliferous material contains at least one metal oxide as catalyst means for the hydrogenation or dehydrogenation of said metalliferous material. According to the inventive method for producing such a metalliferous material, i.e. the actual storage material, the metalliferous material and/or the catalyst means are subjected to a mechanic milling process. (57) Zusammenfassung Die Erfindung betrifft einen metallhaltigen Werkstoff und ein Verfahren zu seiner Herstellung. Der metallhaltige Werkstoff enthält zu seiner Hydrierung oder Dehydrierung wenigstens ein Metalloxid als Katalysationsmittel. Bei dem Verfahren zur Herstellung eines derartigen metallhaltigen Werkstoffs, d.h. des eigentlichen Speicherwerkstoffs, wird der metallhaltige Werkstoff und/oder das Katalysationsmittel einem mechanischen Mahlvorgang unterworfen.		

METALLIFEROUS STORAGE MATERIAL FOR HYDROGEN AND METHOD FOR
PRODUCING SAME

Description

The invention relates to a metalliferrous material and a method of producing the material.

It is first pointed out that, under the term metalliferrous material, atomic metals, metal alloys, intermetallic
5 phases of metals or compound materials as well as corresponding hydrides are to be understood.

It is known that, on the basis of reversible metal hydrides, hydrogen storage devices, so-called hydride storage de-
10 vices, can be formed. The storage device can be charged while heat is released, that is, hydrogen is bound by chemo-sorption and discharged by the application of heat. Hydrogen storage devices can therefore be excellent energy storage devices for mobile and/or stationary applications. They might form in the
15 future a notable storage potential since no noxious emissions are released during the discharge of the hydrogen storage devices.

Very suitable for such hydride storage devices are the so-called mono-crystalline hydrides, which are capable of rapidly storing and releasing the hydrogen. Their manufacture however has been very expensive, so far. Their manufacture, so far, 5 has involved high-energy grinding of elemental components or pre-alloys of nano-crystalline alloys, wherein the grinding procedure can be very long. In a final process step, these nano-crystalline alloys were subjected, depending on the conditions, to a multistage heat treatment under a high hydrogen 10 pressure to be hydrogenated thereby. For many alloys, furthermore, a multiple charging and discharging with hydrogen is necessary to achieve full capacity.

Alternatively, it has been tried to synthesize the respective 15 hydrides by grinding in a hydrogen atmosphere or in a pure chemical way. It has been found, however, that, in this way, the yield of the desired hydrides is smaller and partially additional undesirable phases occur.

20 Furthermore, certain phases could, or respectively can, not be formed with the known conventional methods.

The German patent application No. 197 58 384.6 discloses a method for the manufacture of nano-crystalline metal hydrides 25 with which the manufacture of stable and meta-stable hydrides or hydride-meta-stable alloys is possible with a very high yield of up to 100%. The method described in the mentioned German patent application can be performed with easily controllable limiting conditions and with a relatively small energy 30 consumption.

In order for such a hydrogen storage device to rapidly provide the energy stored therein when needed and to permit rapid charging of the hydrogen storage device, it is desirable

that the reaction speed during hydrating and dehydrating of metals at low temperatures is kept very high that is a very high reaction speed is to be aimed at.

5 To this end, so far, the reaction surface has been increased by reducing the size of the particles / crystals of the materials to be hydrogenated or dehydrogenated as far as this was technically feasible. Other means for increasing the reaction speed included the addition of nickel, platinum or palladium.
10

 The disadvantage of the measures known so far for increasing the reaction speed during the hydrogenation and particularly the dehydrogenation, that is, the delivery of the hydrogen from the hydrogen storage device is that the available
15 speeds are in-sufficient for hydrogen storage devices intended for technical applications.

 It is therefore the object of the present application to
20 provide a metalliferrous material, such as a metal, a metal alloy or an intermetallic phase, compound materials of metals as well as corresponding hydrides with which, during hydrogenation and dehydrogenation, the reaction speeds are so high, that they are technically feasible for use as energy storage devices. A
25 method is to be provided by which the manufacture of a metalliferrous material such as a metal, a metal alloy, an intermetallic phase or a compound material of the materials or corresponding hydrides can be performed in a simple and inexpensive way such that metals manufactured in this way can be used commercially as hydrogen storage devices in a cost-effective manner
30 and with the technically necessary high reaction speed during hydrogenation and dehydrogenation.

The object is solved with regard to the metal-containing material in that it includes a metal oxide as catalyst for the hydrogenation or dehydrogenation thereof.

5 In accordance with the invention, the fact is utilized that, in comparison with pure metals, metal oxides are brittle, whereby a smaller particle size and a homogeneous distribution of the metal oxide in the material is achieved. As a result, the reaction kinetics are substantially increased in comparison
10 with metallic catalysts. Another advantage is that the metal oxides are available as catalysts generally at much lower prices than metals or respectively, metal alloys so that also the aim of commercial utilization at reasonable costs for the metalliferous materials according to the invention can be
15 achieved.

Basically, the metal oxide is an oxide of atomic metals such as the oxide of the metals Mg, Al, Ca, Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Y, Zr, Nb, Ce, Mo, Sn, La, Hf, Ta, W.
20 In accordance with an advantageous embodiment of the invention, the metal oxide may also consist of mixed oxides of the metals, particularly of the metals listed earlier or of mixtures of the metal oxides.

25 Advantageously, the metal oxide or metal oxides may be formed by rare earth metals or metal oxides or mixtures of rare earth metals.

In an advantageous embodiment of the invention, the metal
30 has a nano-crystalline structure, wherein, equally advantageously, also the catalyst has a nano-crystalline structure. If the metal and/or the catalyst have a crystalline structure, the reaction surface and, consequently, the reaction speed of

the hydrogenation or, respectively, the dehydrogenation of the metalliferrous material are increased.

The method according to the invention for the manufacture
5 of such a metalliferrous material is characterized in that the metalliferrous material and/or the catalyst are subjected to a mechanical grinding procedure with the object to form, from both components, a powder with an optimized reaction surface of the metalliferrous material as well as a uniform distribution
10 of the catalyst.

The grinding procedure itself may be selected, depending on the metalliferrous material and/or the catalyst, to be differently long so as to achieve the optimal desired reaction
15 surface and an optimal distribution of the catalyst of the metalliferrous material according to the invention. In this connection, it may be advantageous if the metalliferrous material as such is first subjected to the grinding and the catalyst is added subsequently to the grinding process, however the process
20 may be reversed, that is, the catalyst may be first subjected to the grinding followed by the metalliferrous material. Also, these distinguished possible procedures for the grinding are selected depending on the metalliferrous materials and depending on the catalyst to be added.

25

In order to prevent reactions with the ambient gas during the grinding of the metalliferrous material (metal, metal alloy, intermetallic phase, compound material as well as the hydrides thereof) the method is preferably performed under an inert at-
30 mosphere wherein the inert gas is preferably argon.

As already mentioned, the duration of the grinding process for a metalliferrous material (metal, metal alloy, intermetallic phase, compound material as well as the hydrides thereof) and

the catalyst is variably selectable depending on the metalliferous material and the selected catalyst. Preferably, the duration of the grinding process is in the area of 1 to 200 hours.

5

In another type of the method for the manufacture of a metalliferous material, which may be used as electrode material at least for secondary elements, at least one metal oxide is formed on the surface of the electrode material in situ by contact with oxygen from elements of the electrode/material or by direct supply of oxygen. In this way, a catalyzing oxide can be formed in situ from elements of the hydride storage material.

15 Preferably, during performance of the method, the surface of the electrode material is activated chemically and/or mechanically before the oxide is formed, whereby the oxide formation of the metal can be improved.

20 The invention will now be described in detail with reference to various diagrams, which describe the hydrogenation and dehydrogenation behavior as well as other important parameters.

It is shown in:

25

Fig. 1 an x-ray diffraction diagram after a grinding duration of the metalliferous material of one hour and 200 hours,

Fig. 2a a representation of the sorption behavior of the metalliferous material for the representation of the charging temperature and the charging speed with hydrogen;

30

Fig. 2b the sorption behavior of the metalliferous material at another temperature depending on the charging time,

Fig. 2c a pressure curve with magnesium-hydrogen for the representation of a maximal hydrogen content of the metalliferous material,

5

Fig. 3 X-ray diffraction curves showing the catalyst Cr_2O_3 in the hydrated as well as in the dehydrated state and also traces of MgO and Cr ,

10

Fig. 4a-4d a representation of the improvement of the kinetics during the absorption of hydrogen as well as its desorption,

15

Fig. 5 a typical pattern for the charging capacity during the first 30 charge and discharge cycles of an untreated AB_5 hydride alloy,

20

Fig. 6 the representation of an activation after the first 5 cycles to show an insufficient activation,

Fig. 7 a pattern according to Fig. 5 catalyzed however with a metal oxide according to the invention,

25

Fig. 8 a pattern according to Fig. 6 catalyzed however with a metal oxide according to the invention,

30

Fig. 9 a pattern for the discharge capacity achievable with an untreated alloy in the 10. cycle as a function of the discharge currents applied (with respect to lg alloy),

Fig. 10 a representation according to Fig. 9, but in the 30th cycle,

Fig. 11 a corresponding pattern for the same alloy as in Fig. 9, however, catalyzed with a metal oxide according to the invention, and

5 Fig. 12 a corresponding pattern for the same alloy as in Fig. 10, but catalyzed with a metal oxide according to the invention.

The metalliferous material of the invention may comprise
10 various metals, metal alloys, intermetallic phases, compound materials and corresponding hydrides. They form the storage material of the hydrogen storage devices according to the invention. To accelerate the hydrogenation or the dehydrogenation metal oxides are added as catalysts to these metalliforms
15 materials, wherein the metal oxide may also be a mixed oxide, that is, it may include several metal oxides. Metal oxides, or, respectively, mixed oxides may consist for example of Mg, Al, Ca, Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Y, Zr, Nb, Mo, Sn, Ce, La, Hf, Ta, W or of rare earth. The above listing
20 however is not to be understood in such a way that it represents a limitation of the metal oxides according to the invention to oxides of these metals. Oxides of metals may be for Al_2O_3 , TiO_2 , V_2O_5 , Cr_2O_3 , Fe_2O_3 , Fe_2O_4 , CuO , Nb_2O_5 , MoO , MoO_2 , etc. The catalyst may also have a nano-crystalline structure.

25

A method for the manufacture of a metalliferous material according to the invention will be described on the basis of an example. In the description reference is made to the figures.

30

Examples:MgH₂ + 5Cr₂O₃

- 5 Experimental particulars: 30.7g MgH₂ and 9.3g Cr₂O₃ were placed into a 250 ml grinding container of steel. 400g steel balls (ball diameter 10mm, ratio powder : balls = 1:10) were added. The powder was subjected to a mechanical high-energy grinding process in a planetary ball mill of the type "Fritsch Pulver-
- 10 isette 5". The grinding process was performed under an argon atmosphere for all together 200 hours. During and after the grinding process small amounts of powder were removed for an X-ray structure analysis. Fig. 1 shows the x-ray diffraction diagrams after a grinding duration of 1 hr and 200 hrs. In ad-
- 15 dition to the MgH₂ also the Cr₂O₃ is detectable after 200 hrs by x-ray structure analysis.

Sorption Behavior: In accordance with Fig. 1, the material can be charged at a temperature of 300°C in 100 sec with 4 wt% hydrogen. At a temperature T = 250°C, a hydrogen content of

20 about 3.6 wt% is reached already after about 50 sec. Also, at T=100°C, a rapid charging is possible. A complete hydrogen discharge is possible at T = 300°C in about 400 sec. At T = 250°C, however, in 1200 sec (see Fig. 2b). In the PCT diagram

25 (Fig. 2c), a maximal hydrogen content of the material of 5 wt% is shown in addition to the pressure level of 1.6 bar, which can be assigned to the system magnesium-hydrogen. Fig. 3 shows x-ray diffraction pictures in which, in addition, to Cr₂O₃, traces of MgO and eventually Cr as inactive phase are shown in

30 the hydrated as well as in the dehydrated state. Furthermore, MgH₂ can be found in the hydrated and Mg can be found in the dehydrated state.

Comparison of magnesium + chromium oxide with pure magnesium:

In accordance with Figs. 4a - 4d a clear improvement of the kinetics during absorption of hydrogen as well as during its desorption is apparent. The samples subjected to the same grinding process have different total capacities of hydrogen. 95 MgH₂ + 5 Cr₂O₃ can store 5 wt % and 100 MgH₂ can store 7.6 wt% hydrogen. This is shown in the PCT diagrams (Fig. 4c). Fig. 4a shows an increase of the absorption speed at T=300°C by the factor 10. During desorption at the same temperature a speed advantage with a factor of 6 is achieved (Fig. 4b). The material can be fully dehydrated at T = 250°C in about 1200 sec, if the catalyst Cr₂O₃ is added (Fig. 4d). Pure MgH₂ cannot be dehydrated at T = 250°C within a reasonable period.

15

With reference to Figs. 5 to 12, it is apparent that the acceleration obtainable in accordance with the invention for storing the hydrogen and for the release from the storage material of the electrode (anode) of the accumulator as well as the manufacturing method according to the invention substantially increases the power density and the current density of the accumulator by use of the electrode material, which has been catalyzed in accordance with the invention in comparison with conventional accumulators. As a result, the accumulators according to the invention are suitable for high power applications, for which, so far, only Ni-Cad elements or cells could be used. Furthermore,

more, storage materials can be used for the electrode whose equilibrium pressure is lower at the application conditions and which form more stable hydrides than those that have been common so far. As a result, lower self-discharge rates are achieved.

The acceleration of the kinetics achieved by the catalysts according to the invention compensates for the loss in thermodynamic drive force

toward a hydrogenation/dehydrogenation of the electrode material, so that, in spite of the greater stability of the hydride, current densities are achieved which are sufficient for the application. The oxide catalyst according to the invention
5 or, respectively, the catalyst additions can be manufactured or provided at substantially lower costs than the metals used so far. The activation procedure for the electrode material used so far is eliminated with the manufacture of the metalliferous electrode material according
10 ing to the invention.

It is apparent from figures 5 to 12 that the charging and discharging behavior of the electrode material according to the invention provides for extraordinarily large advantages and improvements when compared with the corresponding behavior of the
15 conventional electrode materials.

It is basically possible to use the electrode material according to the invention also for electrodes of non-rechargeable primary elements or cells, which however could be
20 regenerated.

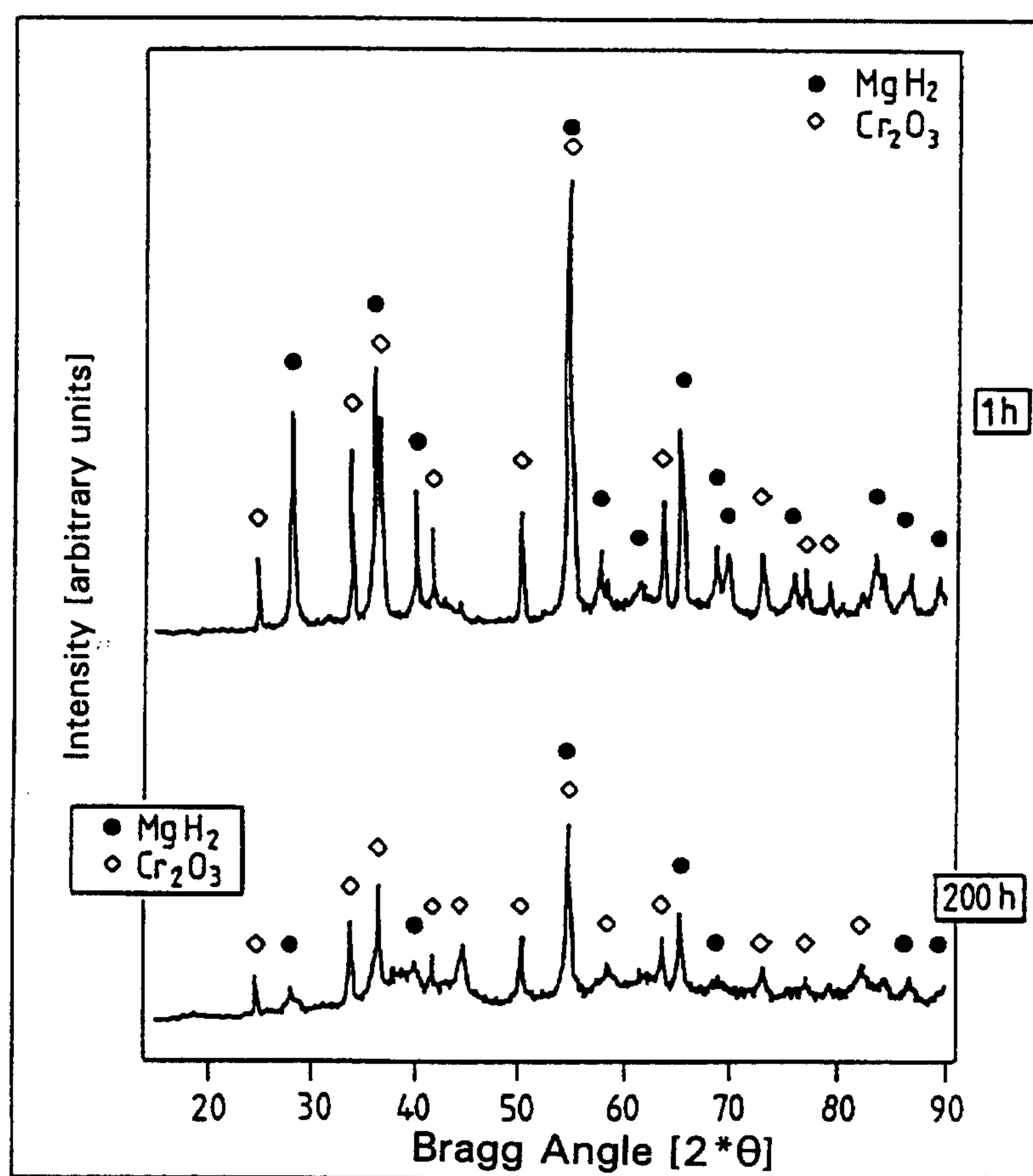
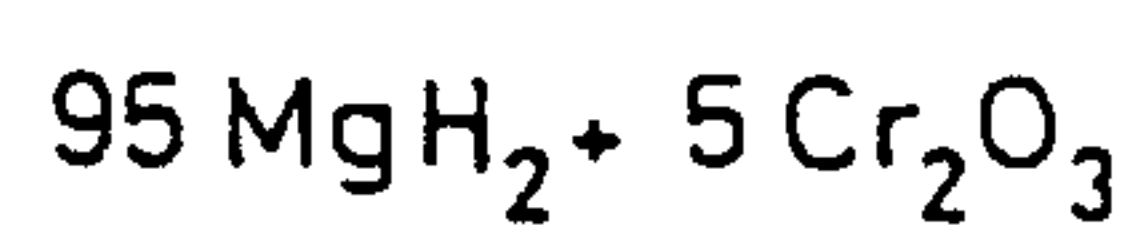
The embodiments of the invention in which an exclusive property or privilege is claimed are defined as follows:

1. A metalliferous storage material for hydrogen including a metal oxide as catalyst for the hydrogenation or dehydrogenation of the metalliferous storage material, said metal oxide of said catalyst including as the metal at least one selected from the group consisting of Ti, V, Cr, Mn, Fe and Nb, wherein said metal oxide of said catalyst has a nanocrystalline structure.
2. A metalliferous storage material according to claim 1, wherein said metal oxide is a mixed oxide.
3. A metalliferous storage material according to claim 1, wherein said metal has a nanocrystalline structure.
4. A method of producing a metalliferous storage material according to claim 1, said method comprising the step of subjecting the metalliferous material and said catalyst to a mechanical grinding process.
5. A method according to claim 4, wherein said grinding process is performed for a predetermined time.
6. A method according to claim 4, wherein said metalliferous material is first subjected to said grinding process and said catalyst is subsequently added to said grinding process.
7. A method according to claim 4, wherein said catalyst is first subjected to said grinding process and said metalliferous material is subsequently added to said grinding process.
8. A method according to claim 4, wherein said grinding process is performed in a protective inert gas atmosphere.

9. A method according to claim 8, wherein said inert gas is argon.
10. A method according to claim 8, wherein the duration of said mechanical grinding process is 1 to 200 hours.
11. A method of manufacturing an electrode at least for use as a secondary element, said electrode having a surface of Ti, V, Cr, Mn, Fe or Nb, said method comprising the step of exposing said surface to oxygen to form with the elements of said electrode material, in situ, a metal oxide at least on the surface of said electrode thereby providing for a catalyst facilitating the hydrogenation and dehydrogenation of said electrode material.
12. A method according to claim 13, wherein the surface of said electrode material is chemically activated before being exposed to said oxygen for forming said oxide.
13. A method according to claim 13, wherein the surface of said electrode material is mechanically activated before it is exposed to said oxygen to form the oxide.

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Fig. 1



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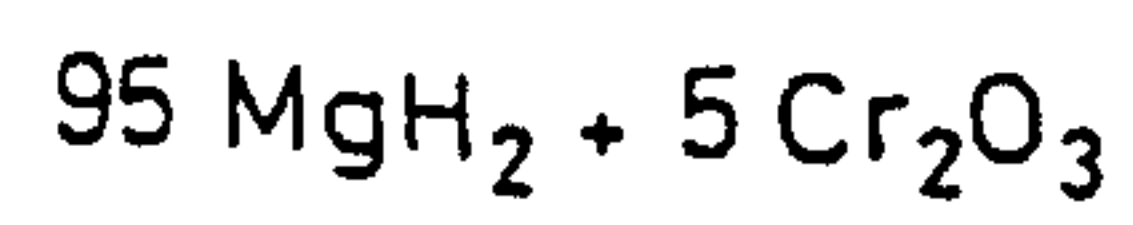


Fig. 2a

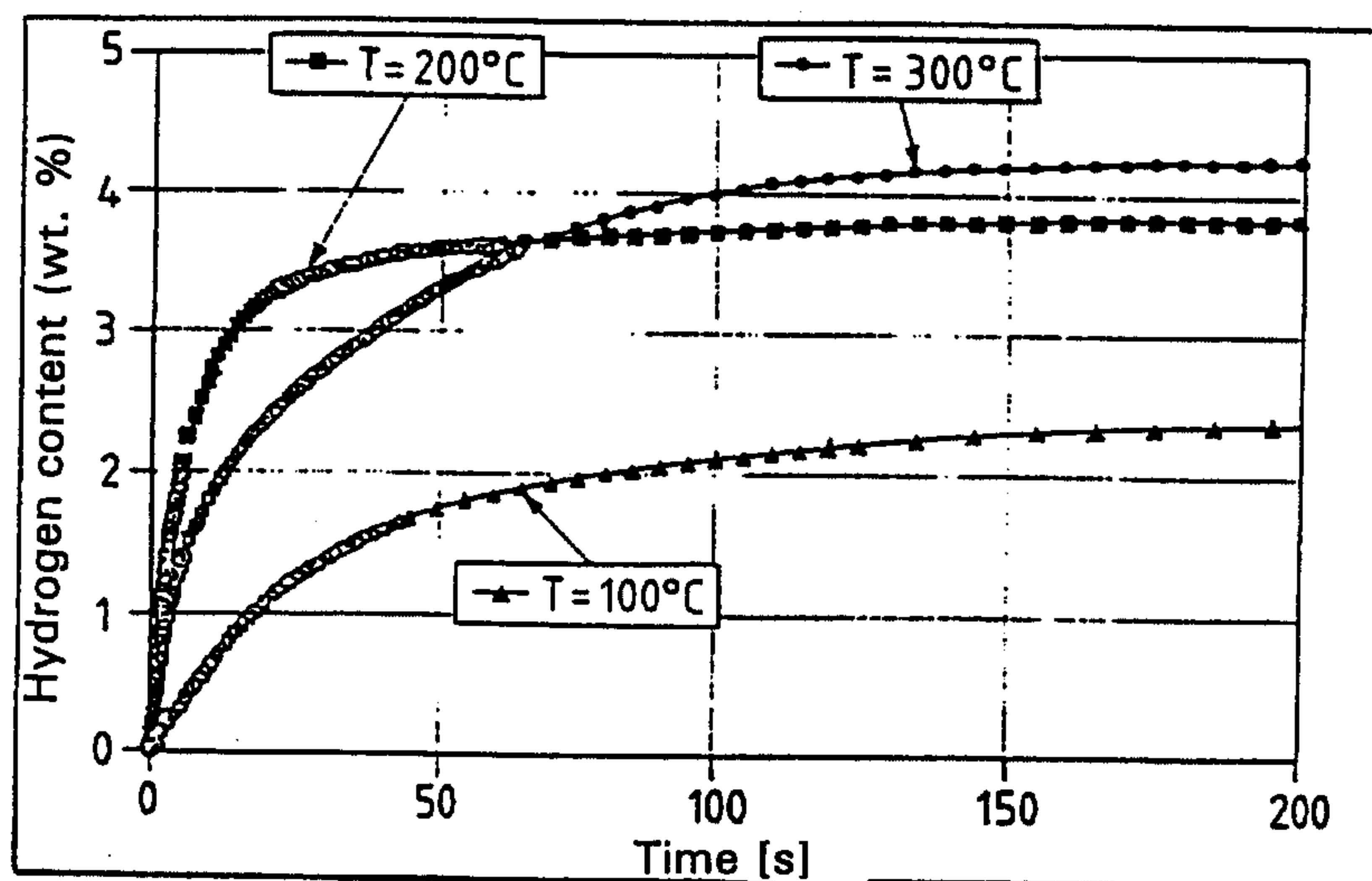
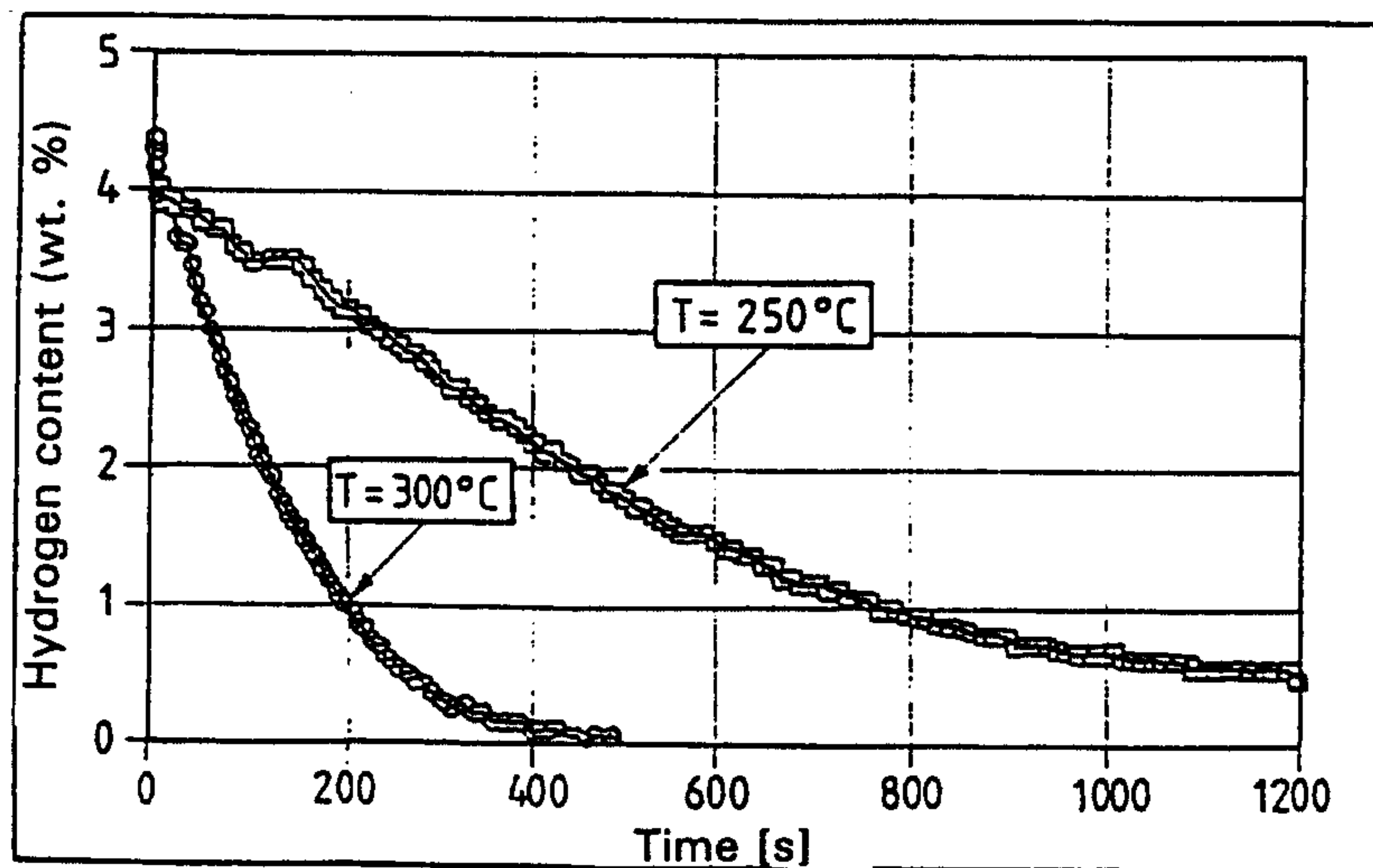
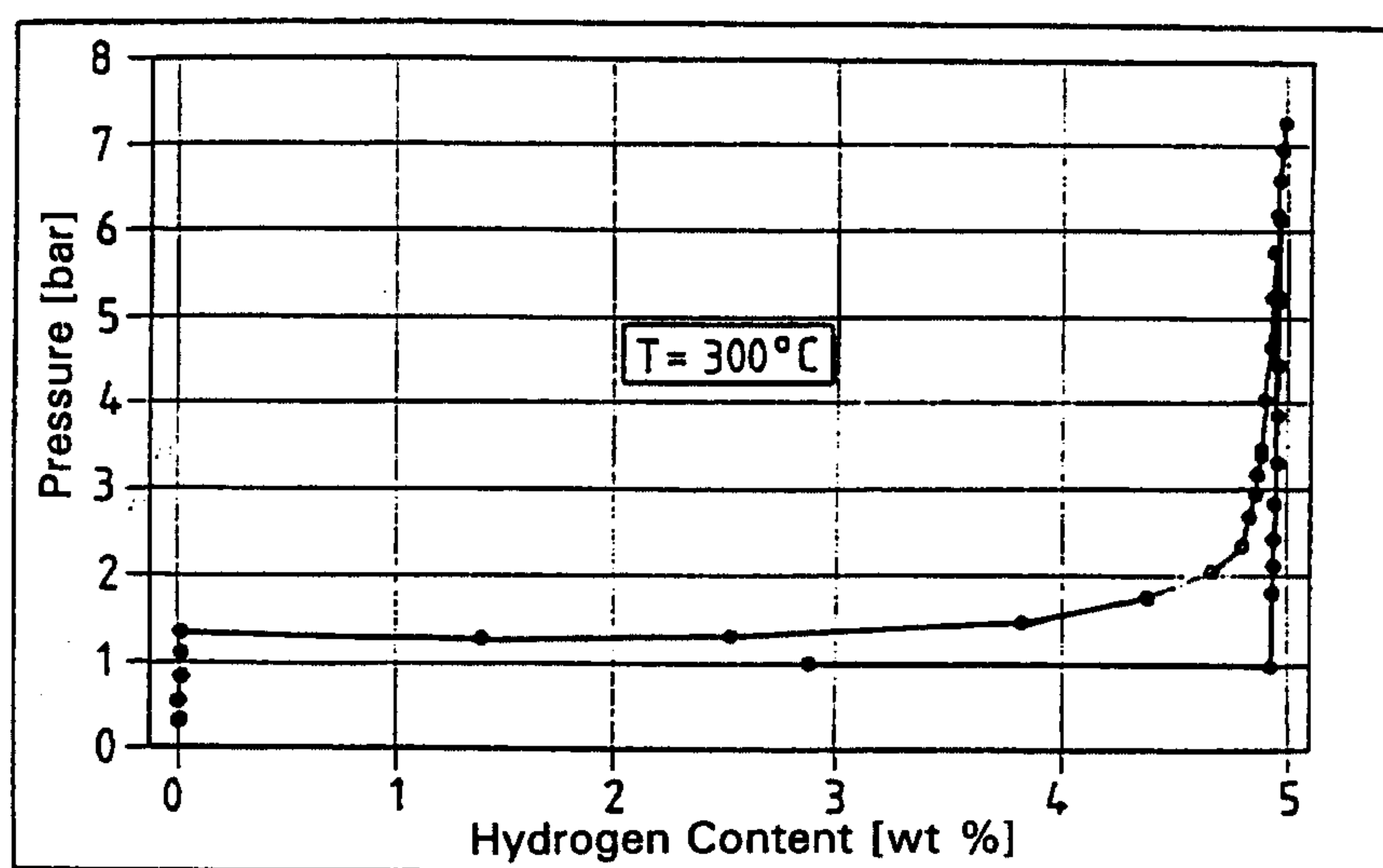


Fig. 2b



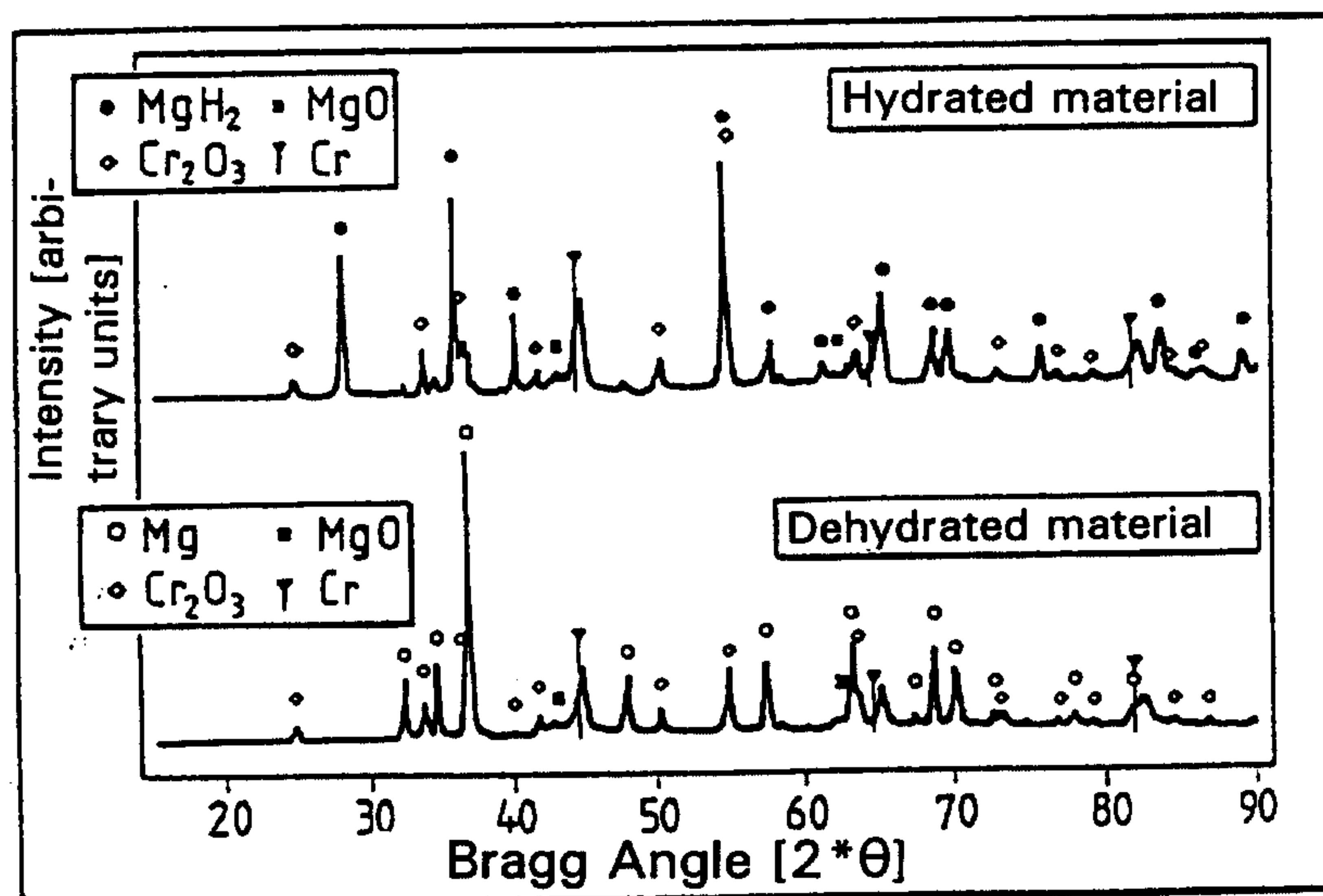
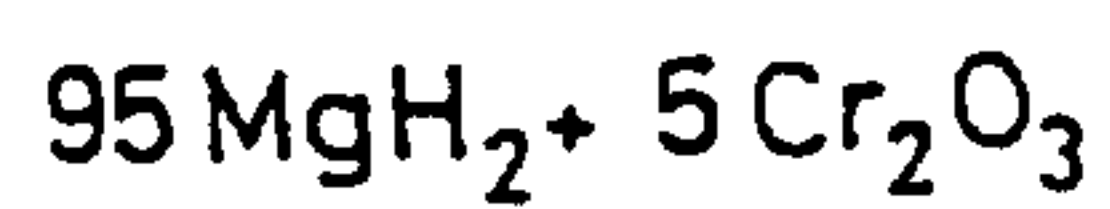
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Fig. 2c

 $95 \text{ MgH}_2 + 5 \text{ Cr}_2\text{O}_3$ 

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Fig. 3



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Fig. 4a

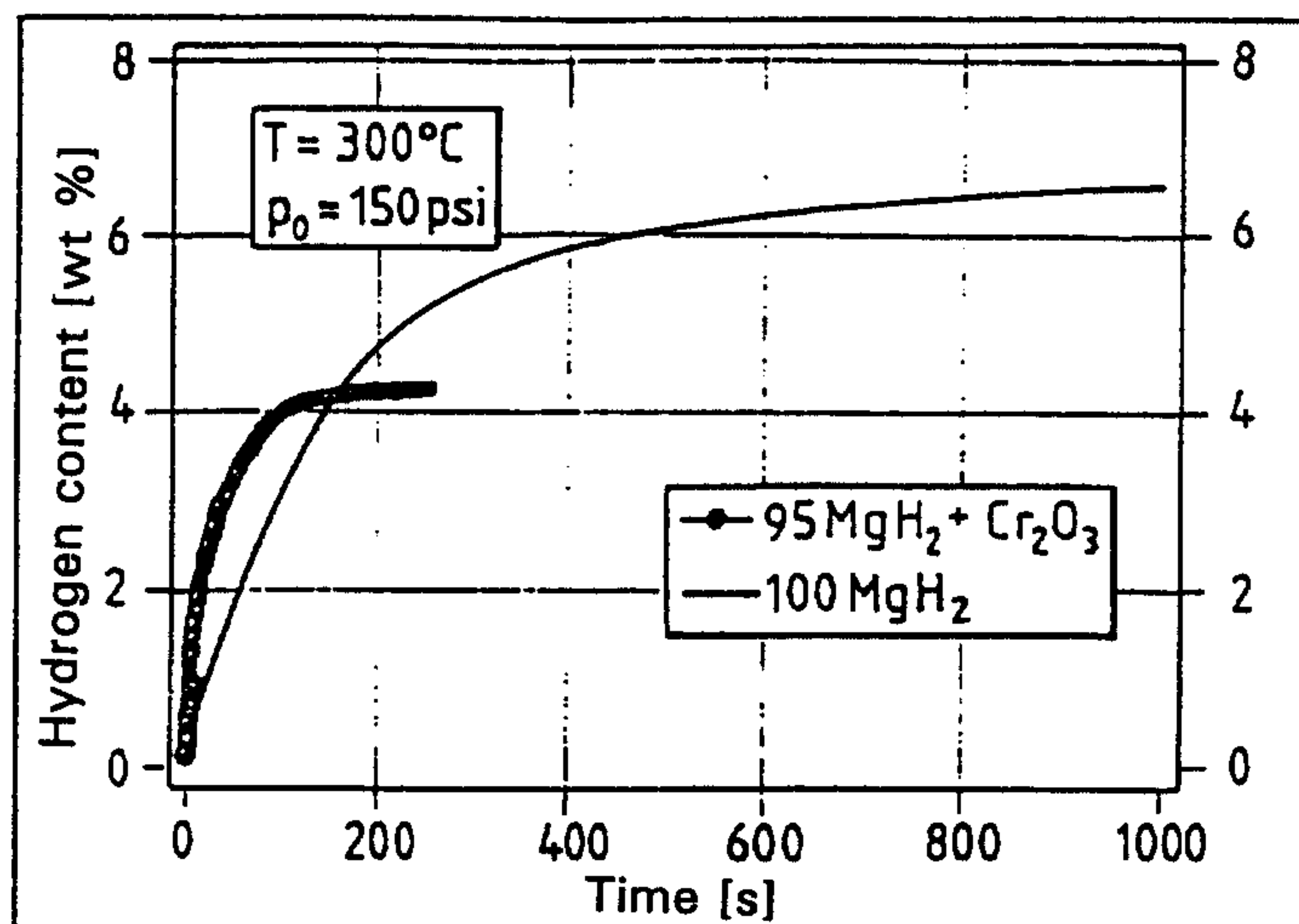
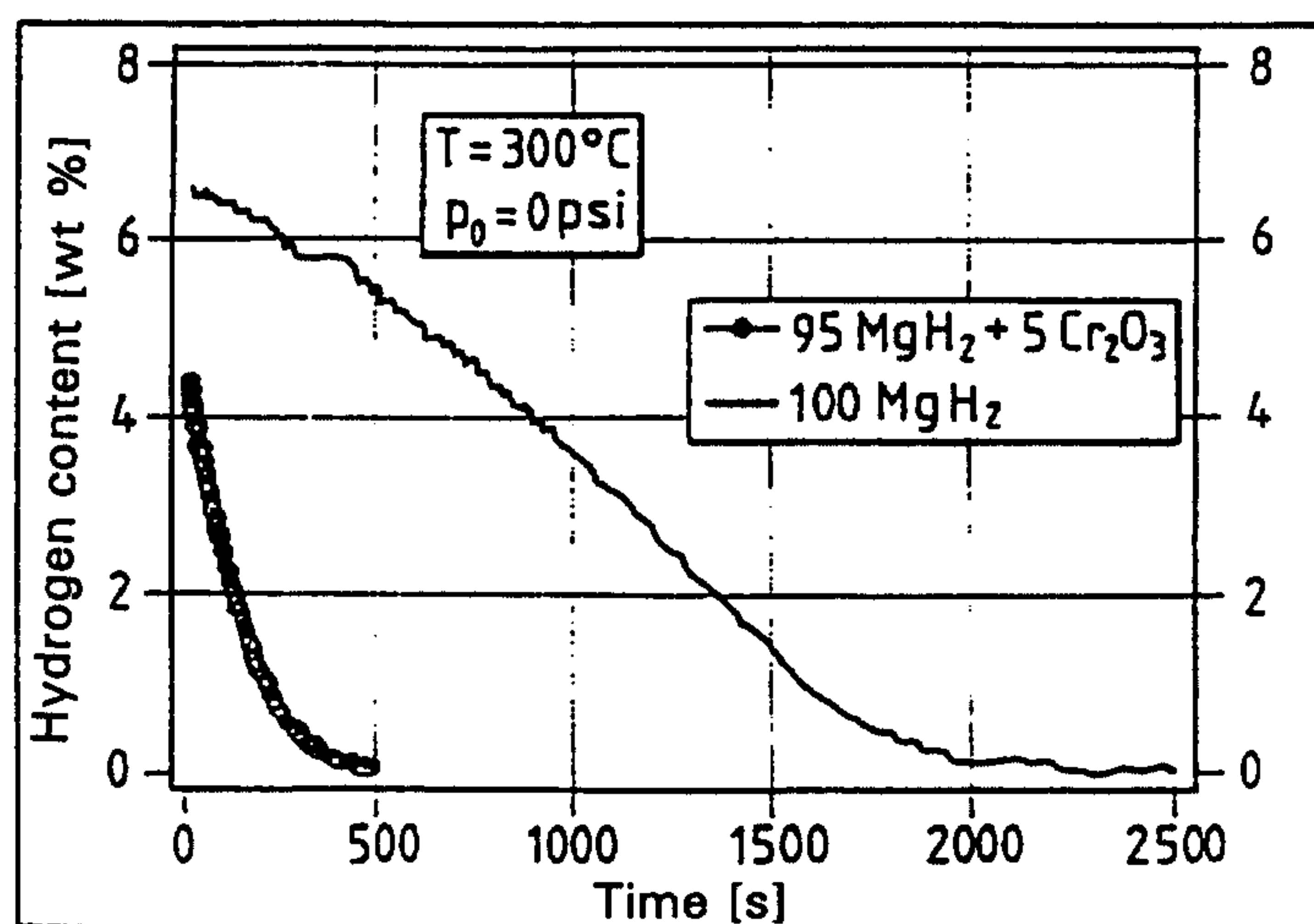


Fig. 4b



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Fig. 4c

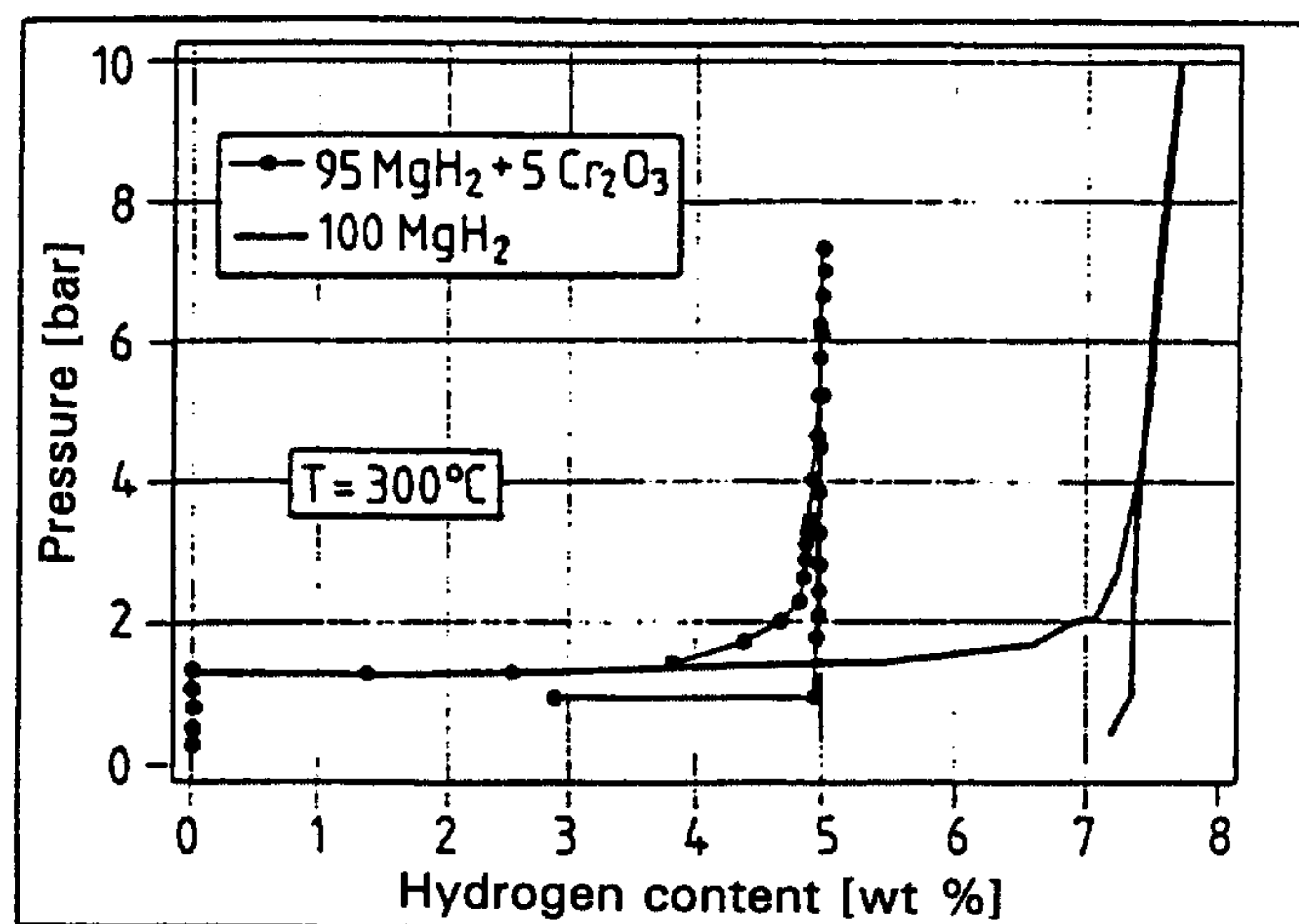
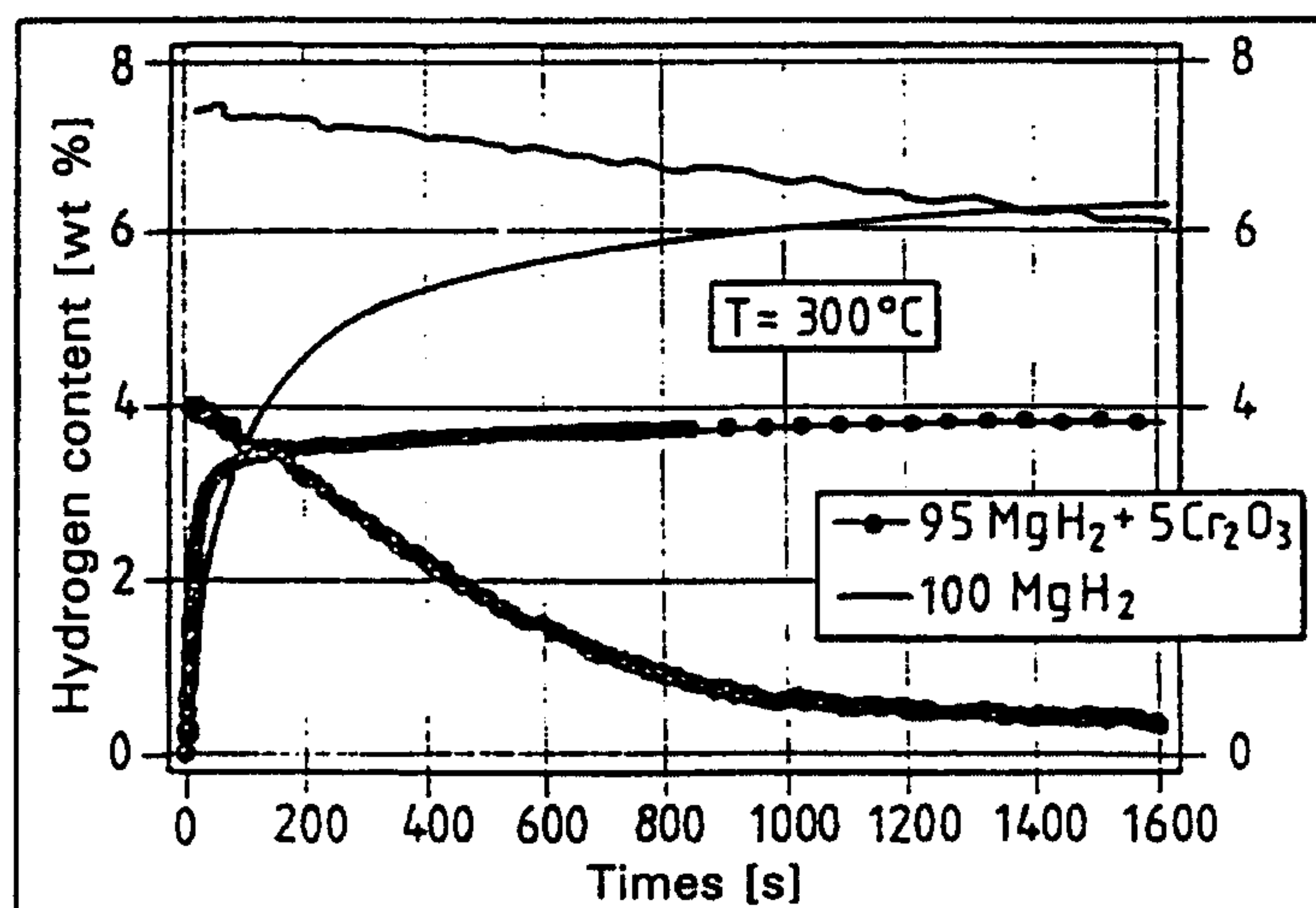


Fig. 4d



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FIG. 5

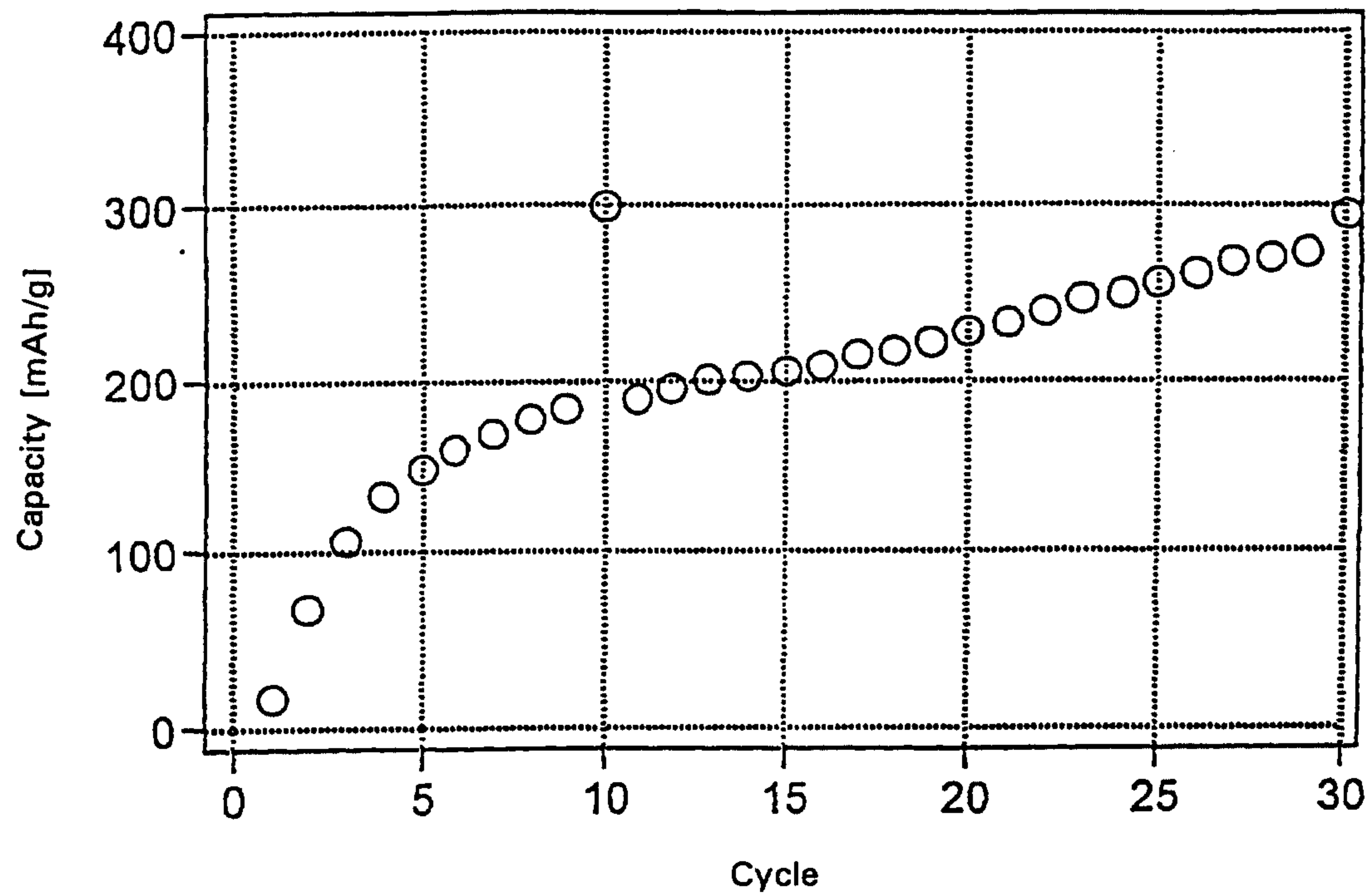
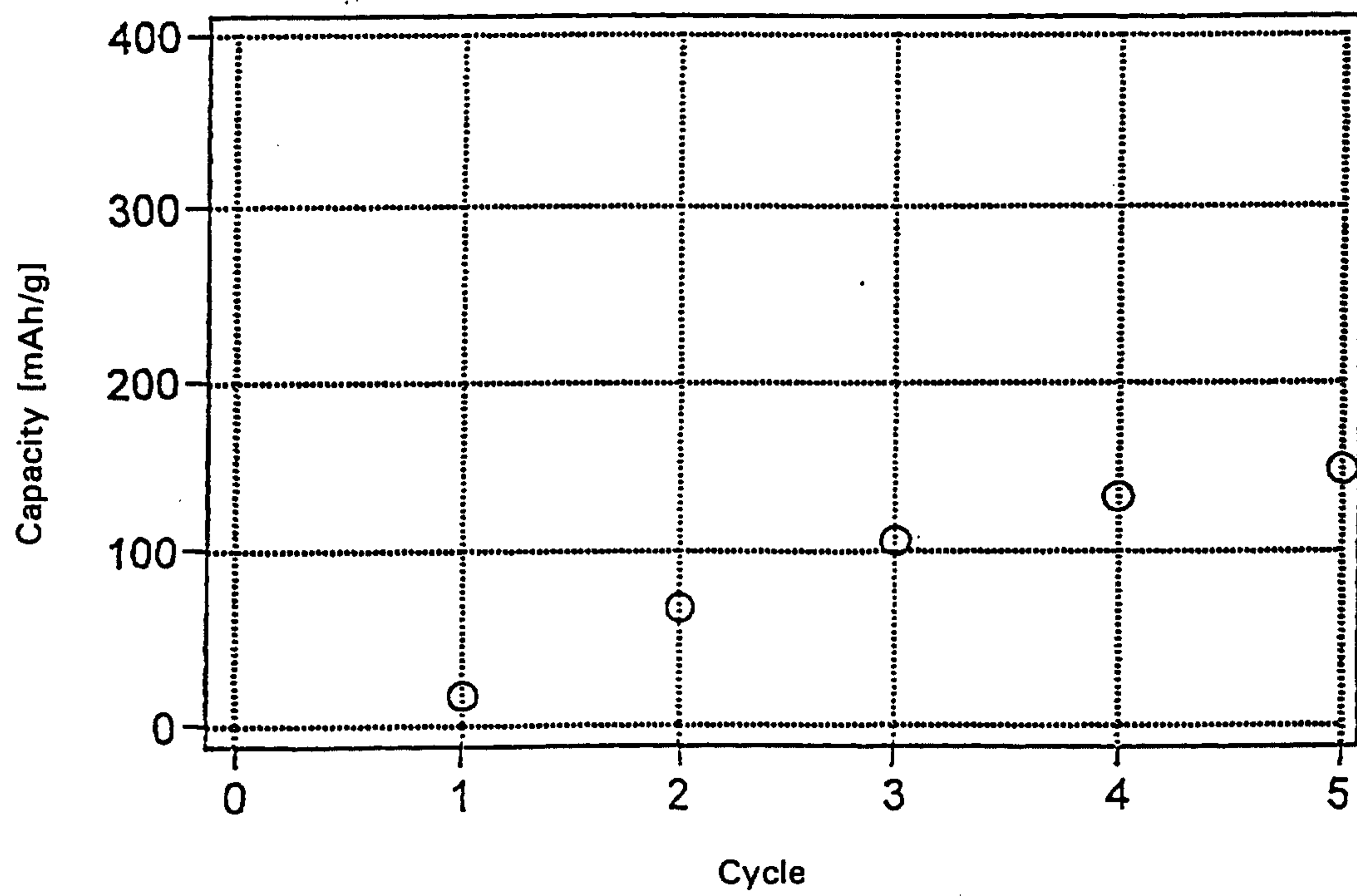


FIG. 6



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FIG. 7

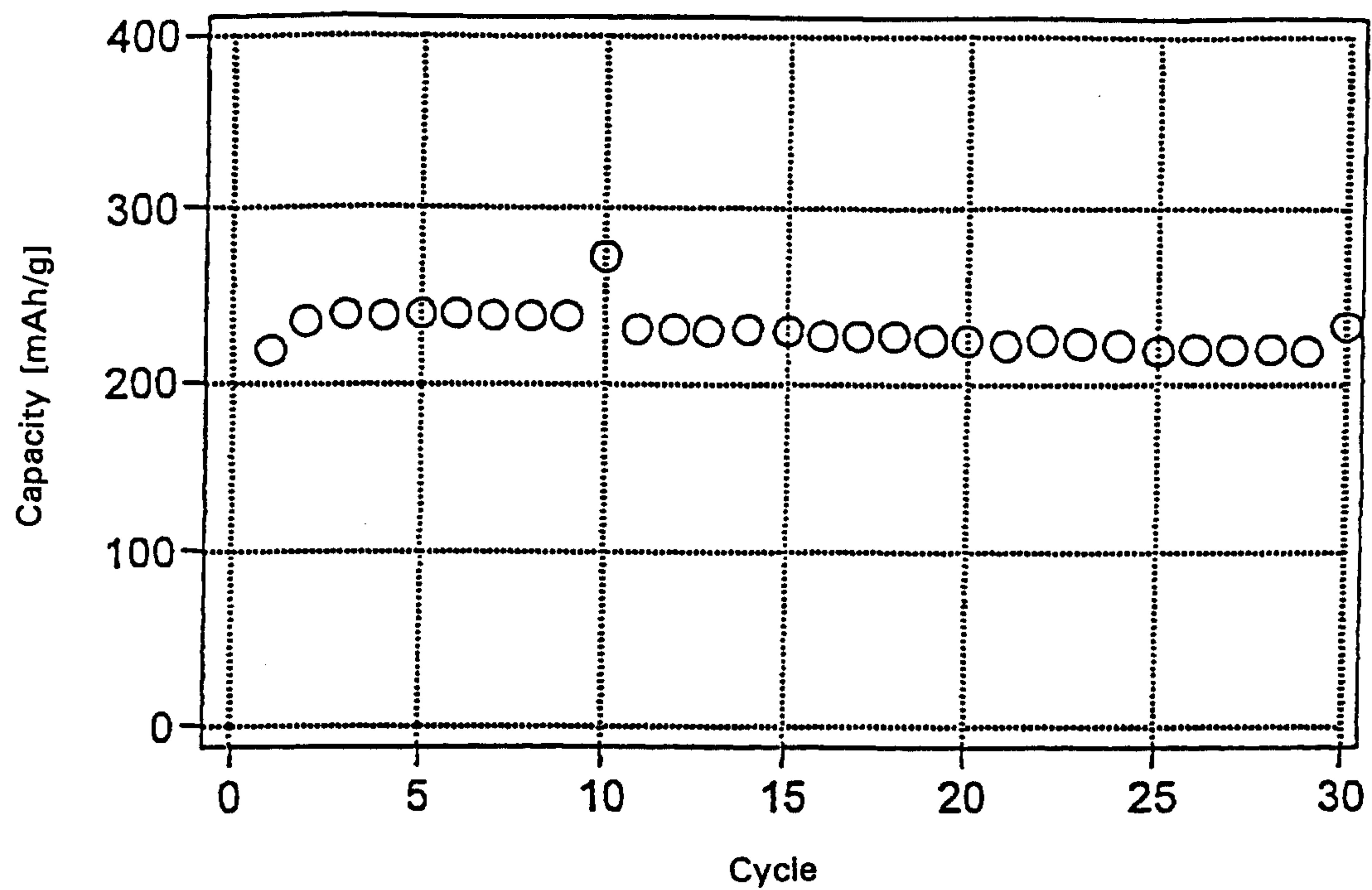
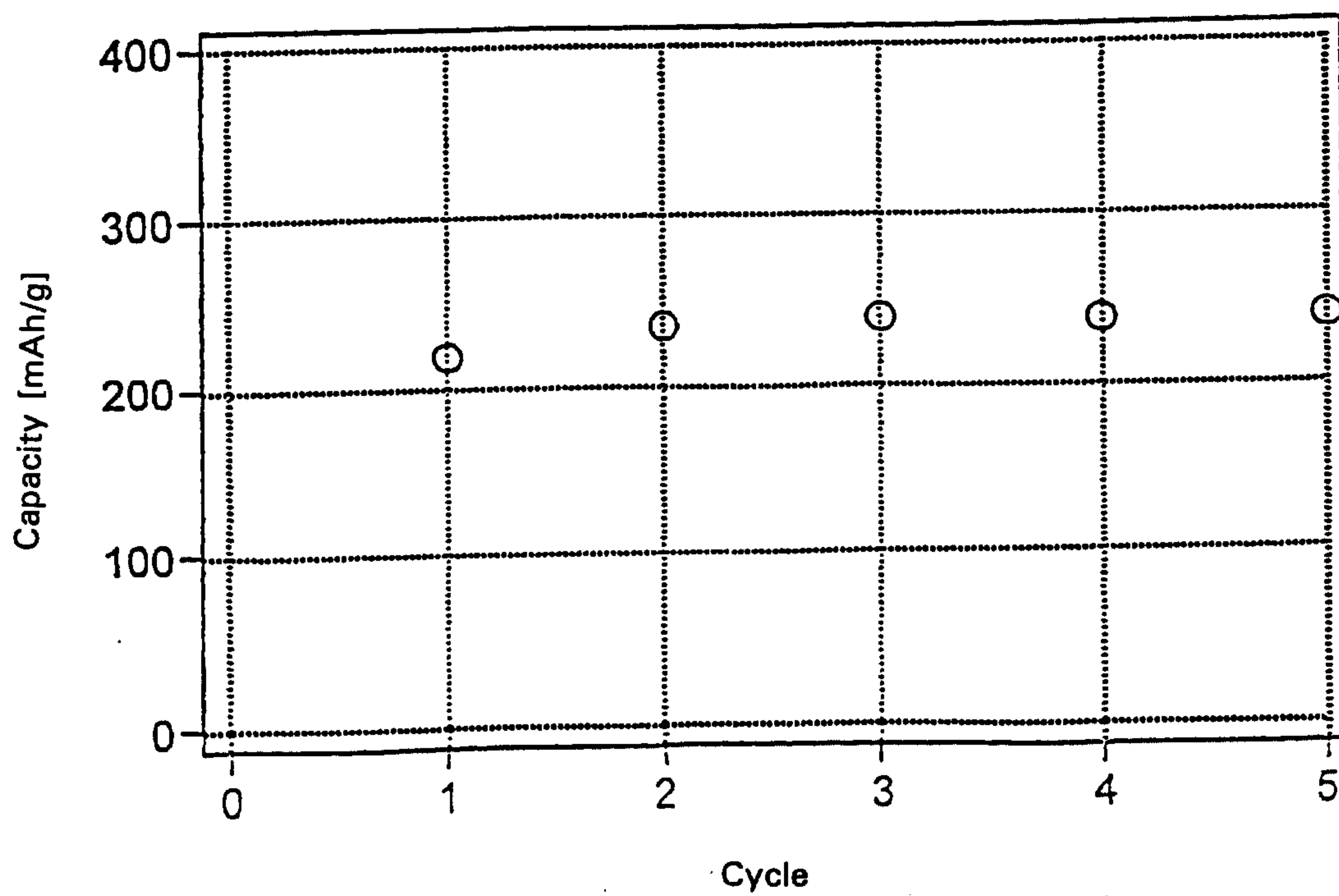


FIG. 8



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Fig. 9

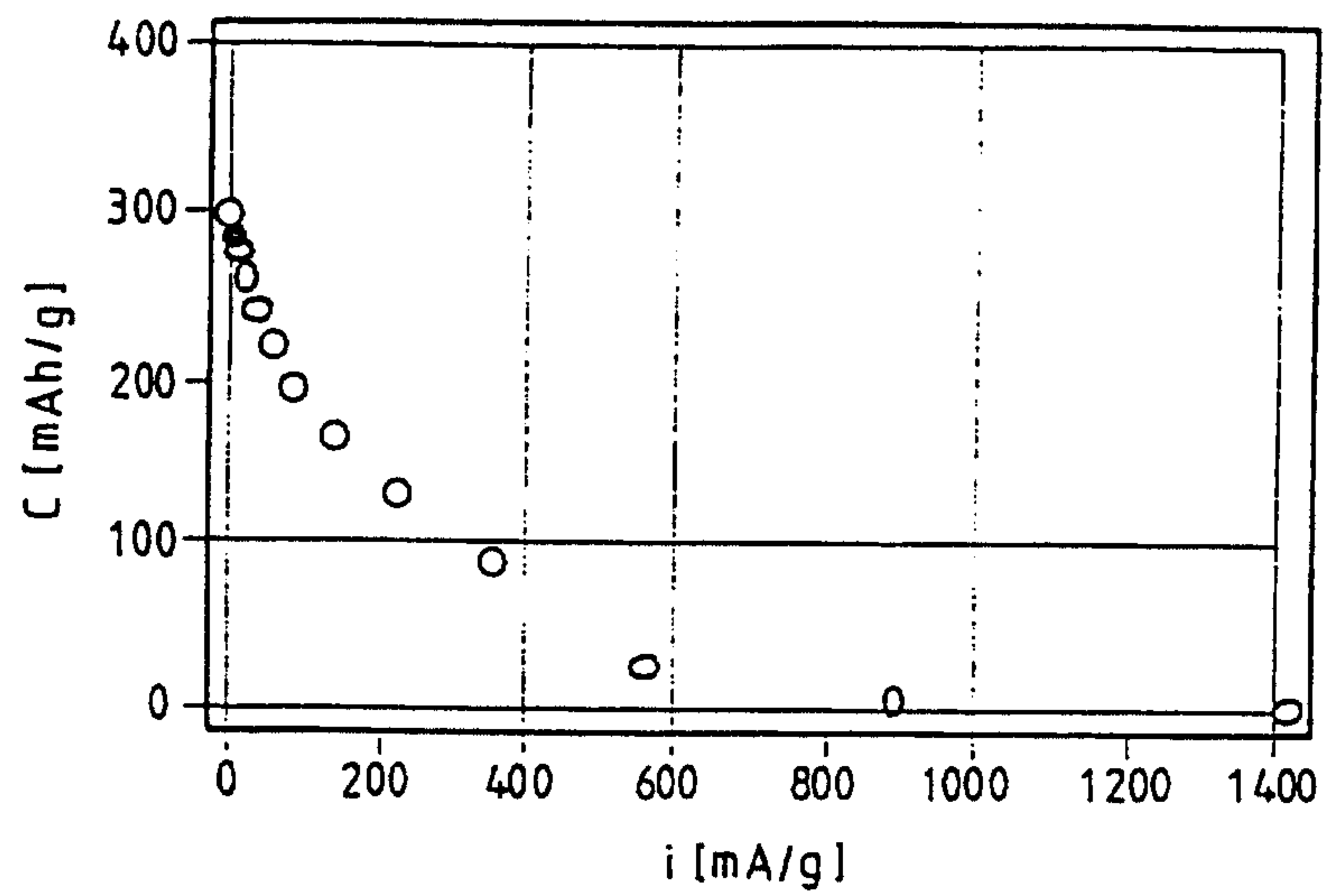
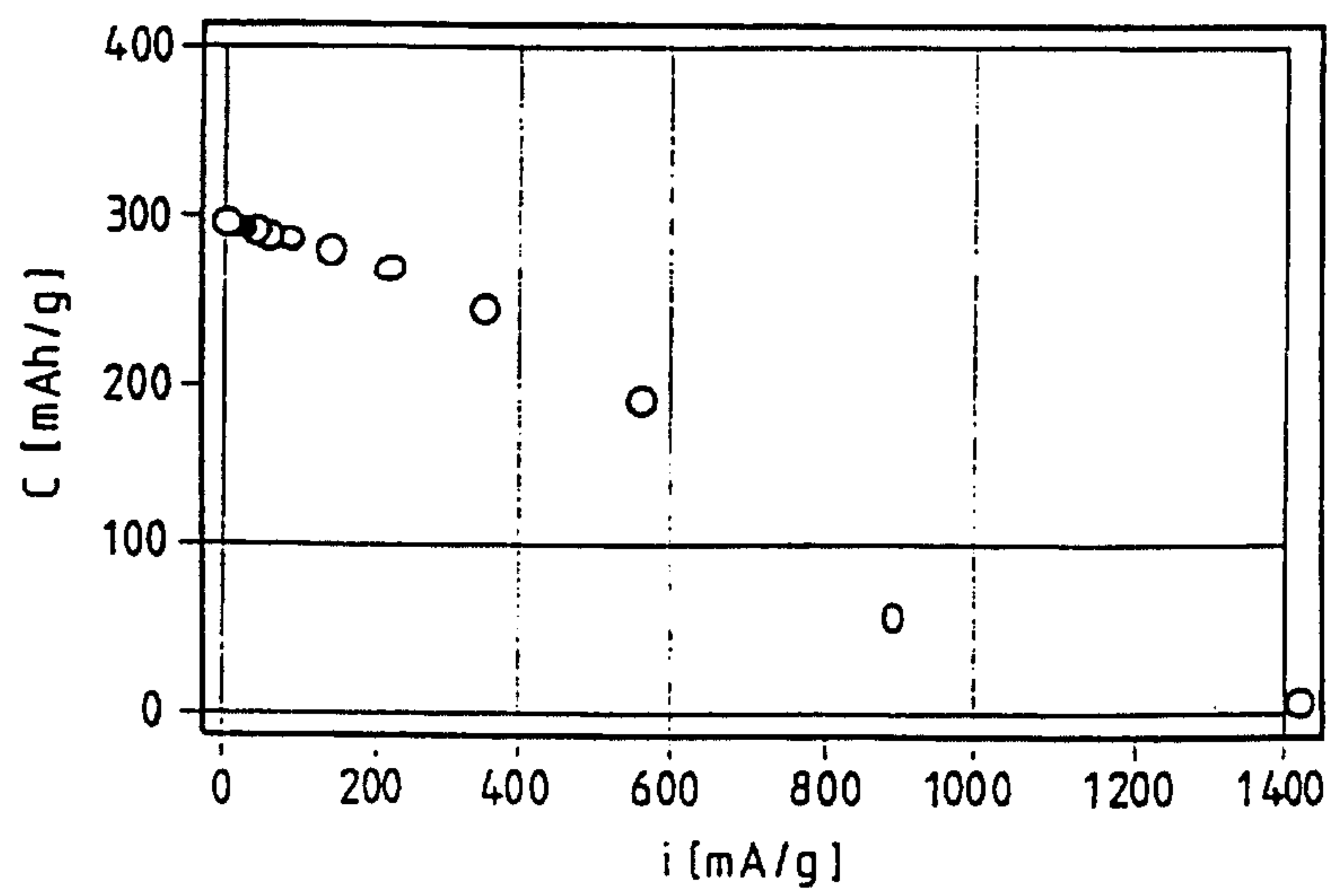


Fig. 10



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Fig. 11

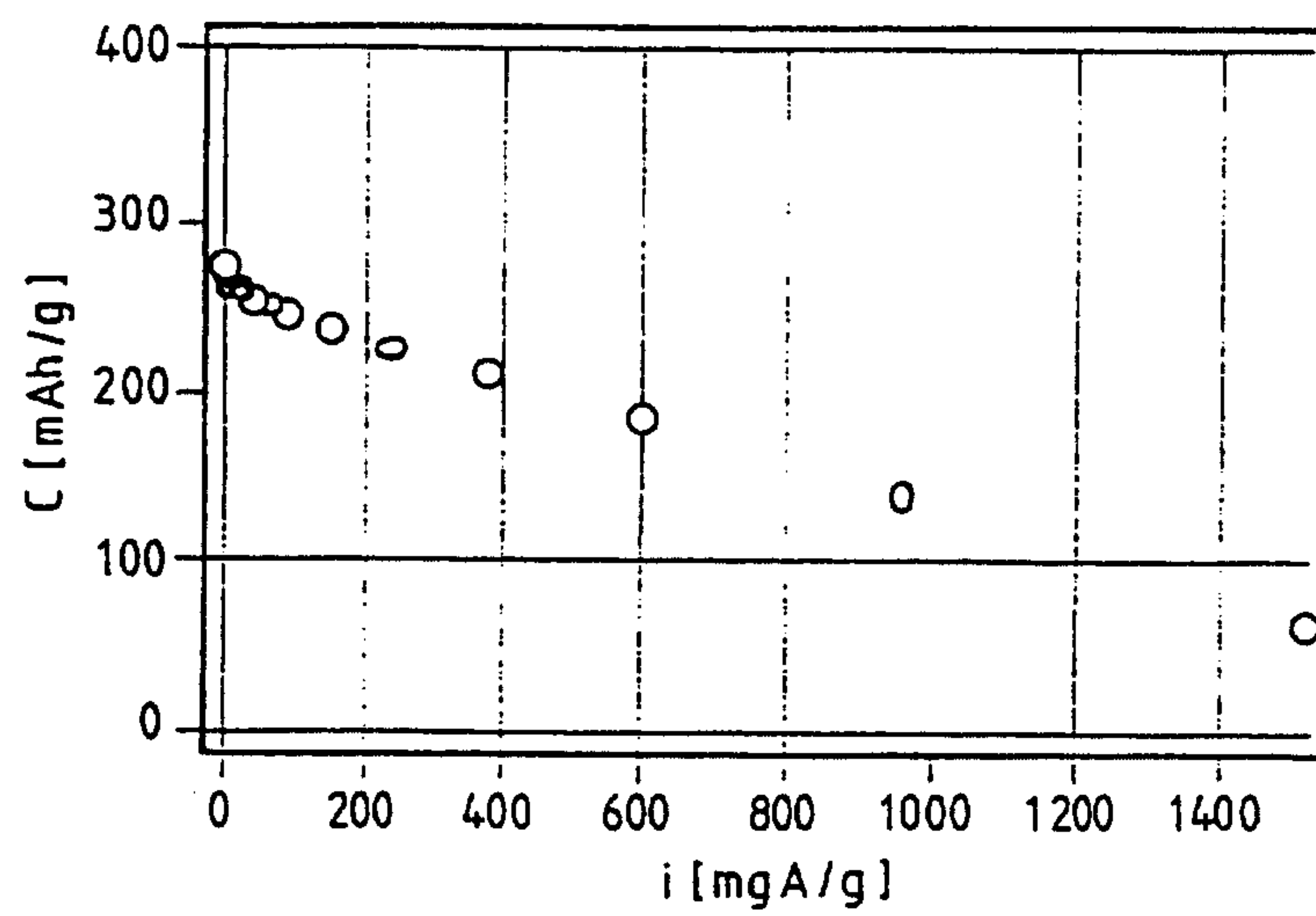


Fig. 12

