



US007343981B2

(12) **United States Patent**
Nappa et al.

(10) **Patent No.:** **US 7,343,981 B2**
(45) **Date of Patent:** **Mar. 18, 2008**

(54) **METHODS USING FLUOROKETONES FOR:
EXTINGUISHING FIRE; PREVENTING FIRE,
AND REDUCING OR ELIMINATING THE
FLAMMABILITY OF A FLAMMABLE
WORKING FLUID**

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(*) Notice: Subject to any disclaimer, the term of this
patent is extended or adjusted under 35
U.S.C. 154(b) by 457 days.

(21) Appl. No.: **10/871,715**

(22) Filed: **Jun. 18, 2004**

(65) **Prior Publication Data**

US 2005/0023007 A1 Feb. 3, 2005

Related U.S. Application Data

(60) Provisional application No. 60/479,559, filed on Jun.
18, 2003.

(51) **Int. Cl.**
A62C 2/00 (2006.01)

(52) **U.S. Cl.** **169/46**; 169/9; 169/11;
169/44; 252/2; 252/8; 568/416

(58) **Field of Classification Search** 169/43,
169/44, 45, 46, 9, 11; 252/2, 8; 568/416,
568/419

See application file for complete search history.

(56) **References Cited**

U.S. PATENT DOCUMENTS

4,420,638 A * 12/1983 Uschold 568/416

5,718,293 A * 2/1998 Flynn et al. 169/45
5,900,185 A 5/1999 Tapscott
5,919,393 A * 7/1999 Flynn et al. 252/2
6,031,011 A 2/2000 Tapscott
6,300,378 B1 10/2001 Tapscott
6,478,979 B1 11/2002 Rivers et al.
6,630,075 B2 * 10/2003 Behr et al. 568/419
7,122,704 B2 * 10/2006 Nappa et al. 252/2

OTHER PUBLICATIONS

Zapevalov, A. Ya. et al., "Synthesis and Reactions of Oxygen-
Containing Organofluorine Compounds. X. Bromoperfluorinated
Carbonyl Compounds and Their Derivatives", Database Caplus
Online Search, Chemical Abstracts Service, Abstract cited (Zhurnal
Organicheskoi Khimii), 1990, pp. 265-272, vol. 26(2), Inst. Khim.,
Sverdlovsk, USSR.
International Search Report.

* cited by examiner

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(57) **ABSTRACT**

The present invention relates to methods of using at least one
fluoroketone selected from monobromoperfluoroketones,
monohydromonobromoperfluoroketones, (perfluoroalkoxy)
monobromoperfluoroketones, (fluoroalkoxy)monobromop-
erfluoroketones, and monochloromonobromoperfluoro-
ketones for i) extinguishing fire by applying to the fire
such a fluoroketone, ii) preventing fire in an air-containing
enclosed area containing combustible materials by introduc-
ing into the area such a fluoroketone and maintaining the
fluoroketone in an amount sufficient to suppress combustion
of combustible materials in the enclosed area, and iii)
reducing or eliminating the flammability of a flammable
working fluid, by mixing between about 0.1 to about 99
percent by weight of such a fluoroketone with the flammable
working fluid.

14 Claims, No Drawings

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**METHODS USING FLUOROKETONES FOR:
EXTINGUISHING FIRE; PREVENTING FIRE,
AND REDUCING OR ELIMINATING THE
FLAMMABILITY OF A FLAMMABLE
WORKING FLUID**

CROSS REFERENCE(S) TO RELATED
APPLICATION(S)

This application claims the priority benefit of U.S. Pro- 10
visional Application 60/479,559, filed Jun. 18, 2003

BACKGROUND OF THE INVENTION

1. Field of the Invention

The present invention relates to methods for extinguishing 15
fire, preventing fire, and reducing or eliminating the flam-
mability of a flammable working fluid using monobromop-
erfluoroketones, monohydrimonobromoperfluoroketones,
(perfluoroalkoxy)monobromoperfluoroketones, (fluoro- 20
alkoxy)monobromoperfluoroketones, and monochlo-
romonobromoperfluoroketones.

2. Description of Related Art

Numerous agents and methods of fire fighting are known 25
and can be selected for a particular fire, depending upon
factors such as its size, location and the type of combustible
materials involved. Halogenated hydrocarbon fire fighting
agents have traditionally been utilized in flooding applica-
tions protecting fixed enclosures (e.g., computer rooms,
storage vaults, telecommunications switching gear rooms,
libraries, document archives, petroleum pipeline pumping 30
stations, and the like), or in streaming applications requiring
rapid extinguishing (e.g., military aircraft, commercial
hand-held extinguishers). Such extinguishing agents are not
only effective but, unlike water, also function as “clean 35
extinguishing agents,” causing little, if any, damage to the
enclosure or its contents.

The most commonly-used halogenated hydrocarbon 40
extinguishing agents have been the bromine-containing
compounds bromotrifluoromethane (CF₃Br, Halon™1301)
and bromochlorodifluoromethane (CF₂ClBr, Halon™1211).
These bromine-containing halocarbons are highly effective
in extinguishing fires and can be dispensed either from 45
portable streaming equipment or from an automatic room
flooding system activated either manually or by some
method of fire detection. However, these compounds have
been linked to ozone depletion. The Montreal Protocol and
its attendant amendments have mandated that Halon™1211
and 1301 production be discontinued.

Thus, there is a need in this field for methods using 50
substitutes or replacements for the commonly-used, bro-
mine-containing fire extinguishing agents. Such substitutes
used in such methods should have a low ozone depletion
potential; should have the ability to extinguish, control, and 55
prevent fires, e.g., Class A (trash, wood, or paper), Class B
(flammable liquids or greases), and/or Class C (electrical
equipment) fires; and should be “clean extinguishing
agents,” i.e., be electrically non-conducting, volatile or
gaseous, and leave no residue upon use. Preferably, such
substitutes used in such methods will also be low in toxicity, 60
not form flammable mixtures in air, have acceptable thermal
and chemical stability for use in extinguishing applications,
and have short atmospheric lifetimes and low global warm-
ing potentials.

Various different fluorinated hydrocarbons have been sug- 65
gested for use as fire fighting agents. For example, in U.S.
Pat. No. 6,300,378, Tapscott discloses tropodegradeable

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bromine-containing halocarbon additives to decrease the
flammability of refrigerants, foam blowing agents, solvents,
aerosol propellants, and sterilants. In U.S. Pat. No. 6,478,
979, Rivers et al. disclose the use of perfluorinated ketones
5 in fire extinguishing.

BRIEF SUMMARY OF THE INVENTION

The aforementioned objectives of methods using substi-
tutes or replacements for the commonly-used, bromine-
containing fire extinguishing agents are met by the present
invention which comprises methods for extinguishing fire,
preventing fire, and reducing or eliminating the flammability
of a flammable working fluid using monobromoperfluoroke-
tones, monohydrimonobromoperfluoroketones, (perfluoro-
alkoxy)monobromoperfluoroketones, (fluoroalkoxy)mono-
bromoperfluoroketones, and 15
monochloromonobromoperfluoroketones.

DETAILED DESCRIPTION OF THE
INVENTION

The term “fluoroketones” will be at points herein to refer
collectively to the classes monobromoperfluoroketones,
(perfluoroalkoxy)monobromoperfluoroketones, (fluoro-
alkoxy)monobromoperfluoroketones, monohydrimonobro-
moperfluoroketones, and monochloromonobromoperfluor-
oketones, that may be used in the methods of the present
invention.

The present fluoroketones may be utilized alone, in com-
bination with one another, or in combination with a co-fire-
fighting agent or propellant selected from known fire fight-
ing agents of the classes hydrofluorocarbons, hydrochloro-
fluorocarbons, perfluorocarbons, perfluoroke- 30
tones, perfluoropolyethers, hydrofluoropolyethers, hydro-
fluoroethers, chlorofluorocarbons, bromofluorocarbons, bro-
mochlorofluorocarbons, hydrobromocarbons,
iodofluorocarbons, and hydrobromofluorocarbons. Such co-
agents can be chosen to enhance the fire fighting capabilities
or modify the physical properties (e.g., modify the rate of
introduction by serving as a propellant) of a fire fighting
composition for a particular type (or size or location) of fire
hazard and can preferably be utilized in ratios (of co-agent
to fluoroketone) such that the resulting composition does not
form flammable mixtures in air. Such fire fighting mixtures
may contain from about 10-90% by weight of at least one
fluoroketone and from about 90-10% by weight of at least
one co-agent.

The present fluoroketones may be utilized additionally in
combination with a propellant (e.g., for expelling a liquid
fluoroketone from a sealed vessel), where the propellant is
moderately flammable or flammable, provided that the
resultant composition comprising fluoroketone and such
propellant is non-flammable.

Of particular utility are azeotropic and azeotrope-like
mixtures containing the present fluoroketones and one or
more compounds selected from the group consisting of
perfluoroketones and hydrofluorocarbons. Such mixtures
may provide a fire fighting composition with a lower boiling
point than either constituent of the mixture as well as
provide a constant ratio of the components of the mixture
during discharge.

The present fluoroketones may be solids, liquids, or gases
under ambient conditions, but are preferably utilized for the
present methods of fire preventing and extinguishing in
either the liquid or the gaseous state (or both). Thus,
normally solid compounds are preferably utilized after trans-

formation to liquid and/or gas through melting, sublimation, or dissolution in a liquid co-agent. Such transformation can occur upon exposure of the compound to the heat of a fire.

The present invention includes a method of extinguishing a fire comprising applying to said fire a composition comprising at least one compound selected from the group consisting of monobromoperfluoroketones, monohydrimonobromoperfluoroketones, (perfluoroalkoxy)monobromoperfluoroketones, (fluoroalkoxy)monobromoperfluoroketones, and monochloromonobromoperfluoroketones in an amount sufficient to extinguish said fire.

The extinguishing method of the present invention can be carried out by introducing the composition into an enclosed area surrounding a fire. Any of the known methods of introduction can be utilized provided that appropriate quantities of the composition are metered into the enclosed area at appropriate intervals. For example, a composition can be introduced by streaming, e.g., using conventional portable (or fixed) fire extinguishing equipment; by misting; or by flooding, e.g., by releasing (using appropriate piping, valves, and controls) the composition into an enclosed area surrounding a fire. The composition can optionally be combined with an inert propellant, e.g., nitrogen, argon, decomposition products of glycidyl azide polymers or carbon dioxide, to increase the rate of discharge of the composition from the streaming or flooding equipment utilized. When the composition is to be introduced by streaming or local application, fluoroketones having normal boiling points in the range of from about 40° C. to about 130° C. (especially fluoroketones that are liquid under ambient conditions) are preferably utilized. When the composition is to be introduced by misting, fluoroketones having boiling points in the range of from about 40° C. to about 110° C. are generally preferred. And, when the composition is to be introduced by flooding, fluoroketones having boiling points in the range of from about 40° C. to about 80° C. are generally preferred.

Preferably, the extinguishing process of the present invention involves introducing fluoroketone to a fire or flame in an amount sufficient to extinguish the fire or flame. One skilled in this field will recognize that the amount of fluoroketone needed to extinguish a particular fire will depend upon the nature and extent of the hazard. When the fluoroketone is to be introduced by flooding, cup burner test data is useful in determining the amount or concentration of fluoroketone required to extinguish a particular type and size of fire. The amount of fluoroketone used to extinguish fire is generally an average resulting concentration of between about 1 and about 10 percent by gas volume of fluoroketone.

The method the present invention further includes preventing fire in an air-containing enclosed area containing combustible materials comprising introducing into said area a composition comprising at least one compound selected from the group consisting of monobromoperfluoroketones, monohydrimonobromoperfluoroketones, (perfluoroalkoxy) monobromoperfluoroketones, (fluoroalkoxy)monobromoperfluoroketones, and monochloromonobromoperfluoroketones, and maintaining said composition in an amount sufficient to suppress combustion of combustible materials in the enclosed area.

Thus, the present invention further includes a method of using fluoroketones for preventing a combustible material from igniting. The present method using fluoroketones may prevent fires or deflagration in an air-containing, enclosed area that contains combustible materials of the self-sustaining or non-self-sustaining type. This method comprises the step of introducing into an air-containing, enclosed area a non-flammable fire preventing composition that is essen-

tially gaseous that comprises at least one present fluoroketone, the composition being introduced and maintained in an amount sufficient to prevent combustion of combustible materials in the enclosed area.

For fire prevention, fluoroketones (and any co-agent(s) utilized) can be chosen so as to provide a composition that is essentially gaseous under use conditions. Preferred fluoroketones for this method have normal boiling points in the range of from about 40° C. to about 130° C. The fluoroketone composition is introduced and maintained in an amount sufficient to prevent combustion of combustible materials in the enclosed area. The amount varies with the combustibility of the particular flammable materials present in the enclosed area. Combustibility varies according to chemical composition and according to physical properties such as surface area relative to volume, porosity, etc. The present fluoroketones can be used to eliminate the combustion-sustaining properties of air and to thereby prevent the combustion of flammable materials (e.g., paper, cloth, wood, flammable liquids, and plastic items). The present fluoroketones can be maintained continuously if a threat of fire is always present or can be introduced into an atmosphere as an emergency measure if a threat of fire or deflagration develops.

The present invention further includes a method of reducing or eliminating the flammability of a flammable working fluid, comprising: a) providing an additive comprising at least one compound selected from the group consisting of monobromoperfluoroketones, monohydrimonobromoperfluoroketones, (perfluoroalkoxy)monobromoperfluoroketones, (fluoroalkoxy)monobromoperfluoroketones, and monochloromonobromoperfluoroketones, and b) mixing between about 0.1 to about 99 percent by weight of said additive with said flammable working fluid.

Flammable working fluids may comprise refrigerants (e.g., propane, propylene, difluoromethane (HFC-32), 1,1-difluoroethane (HFC-152a), 1,1,1-trifluoroethane (HFC-143a)), foam blowing agents (e.g., cyclopentane, n-pentane, iso-pentane, n-butane, iso-butane, dimethyl ether, 1,1-difluoroethane (HFC-152a), 1,1-dichloro-1-fluoroethane (HCFC-141b)), solvents (e.g., monochlorotoluenes, benzotrifluorides, volatile methyl siloxanes, terpenes, alcohols, petroleum distillates, hydrocarbons, ethers, esters, ketones), aerosol propellants (e.g., dimethyl ether, 1,1-difluoroethane (HFC-152a)), and/or sterilants (e.g., hydrocarbon epoxides (ethylene oxide)).

This further method of the present invention uses the aforementioned fluoroketones as additives to reduce or eliminate the flammability of normally flammable working fluids. The aforementioned fluoroketones have the characteristics of high effectiveness for flammability reduction, but have short atmospheric lifetimes (on the order of days or weeks) resulting in low ozone depletion potentials and global warming potentials.

The amount of fluoroketone needed will depend on the application, the material whose flammability is to be reduced, and the specific fluoroketone. The fluoroketones will be most useful at concentrations ranging from 1-80% by weight, although the concentration of fluoroketone in the mixtures can range from 0.1-99% by weight. Expedient proportions include 5-40% by weight of fluoroketone for refrigerant mixtures, 5-50% by weight of fluoroketone for foam blowing agent mixtures, 1-99% fluoroketone for solvent mixtures, 5-25% by weight fluoroketone for aerosol propellant mixtures, and 5-40% by weight fluoroketone for sterilant mixtures.

Refrigerants, foam blowing agents, solvents, aerosol propellants, and/or sterilants may be either gases (vapors) or

liquids. In many cases, materials are stored in one form and used in another. For example, foam blowing agents may be stored as a liquid and used as a gas when the foam is actually blown. In some cases, both gaseous and liquid forms are present during use. Thus, refrigerants are present in both vapor and liquid forms during the operation of most refrigerators or heat pumps. In the gas phase, normally flammable refrigerants, foam blowing agents, solvents, aerosol propellants, and/or sterilants containing the flammability reducing fluoroketone will have a reduced flammability due to the presence of the fluoroketone. Of particular importance is the action of the fluoroketone when the refrigerant, foam blowing agent, solvent, aerosol propellant, and/or sterilant is in the liquid state. The present fluoroketones are volatile, though some are more-so and some less-so. Thus, normally flammable liquid refrigerants, foam blowing agents, solvents, aerosol propellants, and sterilants containing these fluoroketones will, upon full or partial evaporation, produce vapors that have lower flammabilities due to the presence of the flammability reducing fluoroketones, which also evaporate. Of particular importance is that release of the fluoroketones when refrigerants, foam blowing agents, solvents, aerosol propellants, and refrigerants evaporate or are otherwise released into an area will aid in reducing flammability of the vapor above the liquid/vapor interface (i.e., combustible liquids) and explosivity of the vapor if released into a volume such as a room.

Monobromoperfluoroketones comprise perfluorinated ketones containing one bromine substituent and may be generally represented by the formula $R^1C(O)R^2$, wherein R^1 is a C_1 - C_5 perfluoroalkyl radical, and R^2 is a C_1 - C_5 monobromoperfluoroalkyl radical. Monobromoperfluoroketones include the known monobromoperfluoroketones, for example: $CBrF_2C(O)CF_3$, $CBrF_2C(O)CF_2CF_3$, $CF_3C(O)CBrFCF_3$, $CF_3C(O)CF_2CBrF_2$, $CBrF_2C(O)CF_2CF_2CF_3$, $CF_3C(O)CBrFCF(CF_3)_2$, $CBrF_2C(O)CF(CF_3)CF_2CF_3$, $CBrF_2CF_2C(O)CF(CF_3)_2$, $CF_3CBrFC(O)CF(CF_3)_2$, $CF_3CF_2C(O)CBr(CF_3)_2$, and $CF_3CF_2C(O)CF(CBrF_2)(CF_3)$.

The new monobromoperfluoroketones $CF_3C(O)CF_2CF_2CBrF_2$, $CF_3CF_2C(O)CF_2CF_2CBrF_2$ and $CF_3C(O)CF_2CF_2CF_2CF_2CBrF_2$, useful in the method of the present invention, may be prepared by bromination of the corresponding monohydroperfluoroketones by the technique of Kolenko and Plashkin in *Izvestiya Akademii Nauk SSSR, Seriya Khimicheskaya*, pages 1648 to 1650 (1977), and by Zapevalov, et al. in *Zhurnal Organicheskoi Khimii*, Vol. 26, pages 265 to 272 (1990). $CF_3C(O)CF_2CF_2CBrF_2$ may be prepared by bromination of monohydroperfluoroketone $CF_3C(O)CF_2CF_2CHF_2$, which may be prepared by isomerization of an epoxide as described by Zapevalov et al. in *Zhurnal Vsesoyuznogo Khimicheskogo Obshchestva im. D. I. Mendeleeva*, Vol. 18, pages 591 to 593 (1973). $CF_3CF_2C(O)CF_2CF_2CBrF_2$ and $CF_3C(O)CF_2CF_2CF_2CF_2CBrF_2$ may be prepared by bromination of monohydroperfluoroketones $CF_3CF_2C(O)CF_2CF_2CHF_2$ and $CF_3C(O)CF_2CF_2CF_2CF_2CHF_2$, respectively, by the technique of Saloutina, et al. in *Zhurnal Organicheskoi Khimii*, Vol. 29, pages 1325 to 1336 (1993).

Preparation of the new monobromoperfluoroketones $CF_3C(O)CF_2CF_2CBrF_2$, $CF_3CF_2C(O)CF_2CF_2CBrF_2$ and $CF_3C(O)CF_2CF_2CF_2CF_2CBrF_2$, useful in the method of the present invention, may be carried out by conversion of the monohydroperfluoroketone terminal C—H bond to a terminal C—Br may be carried out using brominating agents such as elemental bromine, phosphorous pentabromide, or a

mixture of bromine and phosphorous tribromide. The preferred brominating agent is a mixture of bromine and phosphorous tribromide.

Reaction of a monohydroperfluoroketone and a brominating agent may be carried out under substantially anhydrous conditions in the vapor phase or liquid phase in a container fabricated from materials of construction suitable for contact with bromine and hydrogen bromide at temperatures of about 300° C. to 600° C. Examples of such materials of construction include metallic alloys containing a nickel such as, for example, Hastelloy™ C and Hastelloy™ B. The reaction takes place under the autogenous pressures of the reactants at the reaction temperature.

The ratio of the brominating agent to the monohydroperfluoroketones is at least about 1 mole of brominating agent per mole of monohydroperfluoroketone and preferably about 1.3 moles of brominating agent per mole of monohydroperfluoroketone. More than 1.7 moles of brominating agent per mole of monohydroperfluoroketone provides little benefit.

Brominating the monohydroperfluoroketone may be conducted at temperatures of from about 300° C. to about 600° C. Using the preferred brominating agent, the temperature is preferably conducted from about 300° C. to 350° C. Contact times between the brominating agent and the monohydroperfluoroketone may be from about one hour to about twenty hours.

At the end of the contact period the reaction mixture is cooled and then treated with a reagent to decompose the brominating agent such as sodium sulfite. The monobromoperfluoroketone may be isolated by collecting the organic phase followed by distillation.

The new monobromoperfluoroketones $CF_3CF_2C(O)CBrFCF_2CF_3$, $CF_3CF_2CBrFC(O)CF(CF_3)_2$, $(CF_3)_2CBrC(O)CF(CF_3)_2$, $CF_3CF_2C(O)CBr(CF_3)CF_2CF_3$ and $CF_3CBrFC(O)CF(CF_3)CF_2CF_3$, as well as mixtures of monobromoperfluoroketones $CF_3C(O)CBrFCF_2CF_3$ and $CF_3CBrFC(O)CF_2CF_3$, or $CF_3C(O)CBrFCF_2CF_2CF_2CF_3$ and $CF_3CBrFC(O)CF_2CF_2CF_2CF_3$, or $CF_3CF_2C(O)CBrFCF_2CF_2CF_3$ and $CF_3CF_2CBrFC(O)CF_2CF_2CF_3$, useful in the method of the present invention, may be prepared by reacting perfluoroolefin epoxides, such as the epoxide of perfluoro-2-pentene, perfluoro-2-heptene, or perfluoro-3-heptene, with an alkali metal bromide as described by Saloutina et al. in *Izvestiya Akademii Nauk SSSR, Seriya Khimicheskaya*, pages 1893 to 1896 (1982). The perfluoroolefin epoxides may be prepared by reaction of the perfluoroolefin with an alkali metal hypohalite as described by Kolenko, et al. in *Izvestiya Akademii Nauk SSSR, Seriya Khimicheskaya*, pages 2509-2512 (1979).

The reaction of perfluoroolefin epoxides with alkali metal bromides may be carried out in a polar non-protic solvent such as glycol ethers such as ethylene glycol dimethyl ether, diethylene glycol dimethyl ether, triethylene glycol dimethyl ether, and tetraethylene glycol dimethyl ether, N,N-dimethyl formamide, N,N-dimethyl acetamide, dimethyl sulfolane, dimethylsulfoxide, N-methylpyrrolidinone, and alkane nitriles such as acetonitrile, propionitrile, and butyronitrile. Preferred solvents for contacting perfluoroolefin epoxides with alkali metal bromides are glycol ethers such as ethylene glycol dimethyl ether, diethylene glycol dimethyl ether, triethylene glycol dimethyl ether, and tetraethylene glycol dimethyl ether, and alkane nitriles such as acetonitrile, propionitrile, and butyronitrile.

Alkali metal bromides suitable for opening the perfluoroolefin epoxide ring and formation of a C—Br bond include

lithium bromide, sodium bromide, potassium bromide, and cesium bromide; sodium and lithium bromide are preferred.

The mole ratio of the alkali metal bromide to the perfluoroolefin epoxide is at least about 2:1, preferably about 10:1.

Reaction of alkali metal bromides and perfluoroolefin epoxide may be conducted in the liquid phase under substantially anhydrous conditions at temperatures of from about 10° C. to about 150° C., with contact times of from about 0.5 hour to about thirty-six hours. The pressure under which the reaction occurs is not critical.

At the end of the contact period the reaction mixture may be distilled to isolate the monobromoperfluoroketone.

The new monobromoperfluoroketones $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)_2$, $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}_2\text{CF}_3$, $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CF}_2\text{CF}_2\text{CF}_2\text{CBrF}_2$, $\text{CBrF}_2\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)_2$, $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{CF}_2\text{CF}_3$ and $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)_2$ may be prepared by reacting a monobromoperfluoroacyl fluoride with a perfluoroolefin.

$\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)_2$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CF}=\text{CF}_2$; $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}_2\text{CF}_3$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CF}_2$; $\text{CF}_3\text{CF}_2\text{C}(\text{O})\text{CF}_2\text{CF}_2\text{CF}_2\text{CBrF}_2$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CF}_2$; $\text{CBrF}_2\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)_2$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CF}=\text{CF}_2$; $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{CF}_2\text{CF}_3$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CF}=\text{CF}_3$; and $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)_2$ may be prepared by reacting $\text{CF}_3\text{CBrFC}(\text{O})\text{F}$ with $\text{CF}_3\text{CF}=\text{CF}_2$.

The new monobromoperfluoroketones $\text{CF}_3\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$, $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}_2\text{CF}_2\text{CF}_3$, $\text{CF}_3\text{C}(\text{O})\text{CBr}(\text{CF}_3)\text{CF}_2\text{CF}_3$, $\text{CF}_3\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{CBrFCF}_3$, $\text{CF}_3\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$, useful in the method of the present invention, may be prepared by reacting a perfluoroacyl fluoride with a monobromoperfluoroolefin.

$\text{CF}_3\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$ may be prepared by reacting $\text{CF}_3\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}_2\text{CF}_2\text{CF}_3$ may be prepared by reacting $\text{CBrF}=\text{CF}_2$ with $\text{CF}_3\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{F}$; a mixture of $\text{CF}_3\text{C}(\text{O})\text{CBr}(\text{CF}_3)\text{CF}_2\text{CF}_3$ and $\text{CF}_3\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{CBrFCF}_3$ may be prepared by reacting $\text{CF}_3\text{CBr}=\text{CF}_3$ with $\text{CF}_3\text{C}(\text{O})\text{F}$; and $\text{CF}_3\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$ may be prepared by reacting $\text{CF}_3\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$.

(Perfluoroalkoxy)monobromoperfluoroketones useful in the process of the present invention are of the formula $\text{R}^1\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OR}^F$, wherein R^1 is a C_1 to C_3 monobromoperfluoroalkyl radical, and R^F is a C_1 to C_3 perfluoroalkyl radical, may be obtained by reacting monobromoperfluoroacyl fluorides of the formula $\text{R}^1\text{C}(\text{O})\text{F}$ with perfluorovinyl ethers of the formula $\text{CF}_2=\text{CFOR}^F$. Representative new (perfluoroalkoxy)monobromoperfluoroketones useful in the process of the present invention include $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$, $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$, $\text{CBrF}_2\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$, $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$, $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$, $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_2\text{C}_2\text{F}_5$, $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_2\text{C}_2\text{F}_5$, $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}(\text{CF}_3)_2$, $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}(\text{CF}_3)_2$, $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}(\text{CF}_3)_2$, $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$, and $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$.

$\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_3$; $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_3$; $\text{CBrF}_2\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_3$; $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$ may be pre-

pared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_2\text{F}_5$; $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_2\text{F}_5$; $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_2\text{C}_2\text{F}_5$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_2\text{C}_2\text{F}_5$; $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_2\text{C}_2\text{F}_5$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_2\text{C}_2\text{F}_5$; $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}(\text{CF}_3)_2$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}(\text{CF}_3)_2$; $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}(\text{CF}_3)_2$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}(\text{CF}_3)_2$; $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$ may be prepared by reacting $\text{CF}_3\text{CBrFC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_3$; and $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$ may be prepared by reacting $\text{CF}_3\text{CBrFC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_3$.

(Perfluoroalkoxy)monobromoperfluoroketones of the formula $\text{R}^1\text{C}(\text{O})\text{CF}(\text{CF}_3)\text{OR}^F$ may also be obtained by reacting perfluoroalkoxyperfluoroacyl fluorides of the formula $\text{R}^F\text{OCF}(\text{CF}_3)\text{C}(\text{O})\text{F}$ with a monobromoperfluoroolefin. Representative (perfluoroalkoxy)monobromoperfluoroketones useful in the process of the present invention include $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$, $(\text{CF}_3)_2\text{CBrC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$, $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$, $(\text{CF}_3)_2\text{CBrC}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$, $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_2\text{C}_2\text{F}_5$, and $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}(\text{CF}_3)_2$.

$\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$ may be prepared by reacting $\text{CF}_3\text{OC}(\text{CF}_3)\text{FC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CBrF}$; $(\text{CF}_3)_2\text{CBrC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_3$ may be prepared by reacting $\text{CF}_3\text{OC}(\text{CF}_3)\text{FC}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$ may be prepared by reacting $\text{C}_2\text{F}_5\text{OC}(\text{CF}_3)\text{FC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CBrF}$; $(\text{CF}_3)_2\text{CBrC}(\text{O})\text{CF}(\text{CF}_3)\text{OC}_2\text{F}_5$ may be prepared by reacting $\text{C}_2\text{F}_5\text{OC}(\text{CF}_3)\text{FC}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}_2\text{C}_2\text{F}_5$ may be prepared by reacting $\text{C}_2\text{F}_5\text{CF}_2\text{OC}(\text{CF}_3)\text{FC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CBrF}$; and $\text{CF}_3\text{CBrFC}(\text{O})\text{CF}(\text{CF}_3)\text{OCF}(\text{CF}_3)_2$ may be prepared by reacting $(\text{CF}_3)_2\text{CFOCF}(\text{CF}_3)\text{FC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CBrF}$.

(Fluoroalkoxy)monobromoperfluoroketones useful in the process of the present invention are of the formula $\text{R}^1\text{C}(\text{O})\text{CX}(\text{CF}_3)\text{OR}^2$, wherein X is H or F, R^1 is a C_1 , C_2 , or C_3 bromoperfluoroalkyl radical, and R^2 is a C_1 to C_3 alkyl or fluoroalkyl radical, may be prepared by reacting monobromoperfluoroacyl fluorides of the formula $\text{R}^1\text{C}(\text{O})\text{F}$ with hydrofluorovinyl ethers of the formula $\text{CF}_2=\text{CXOR}^2$. Representative new (fluoroalkoxy)monobromoperfluoroketones include $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{OCF}_2\text{CHF}_2)\text{CF}_3$, $\text{CBrF}_2\text{C}(\text{O})\text{CH}(\text{OCF}_2\text{CHF}_2)\text{CF}_3$, $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{OCH}_3)\text{CF}_3$, and $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_2\text{OCH}_3)\text{CF}_3$.

$\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{OCF}_2\text{CHF}_2)\text{CF}_3$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCF}_2\text{CHF}_2$; $\text{CBrF}_2\text{C}(\text{O})\text{CH}(\text{OCF}_2\text{CHF}_2)\text{CF}_3$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CHOCF}_2\text{CHF}_2$; and $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{OCH}_3)\text{CF}_3$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFOCH}_3$.

Another (fluoroalkoxy)monobromoperfluoroketone useful in the process of the present invention includes $\text{CBrF}_2\text{C}(\text{O})\text{CF}(\text{CF}_2\text{OCH}_3)\text{CF}_3$, prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CF}=\text{CFOCH}_3$.

The reaction of fluoroacyl fluorides with fluoroolefins is described by Fawcett, et al. in U. S. Pat. No. 3,185,734 and *Journal of the American Chemical Society*, Vol. 84, pages 4285 to 4288 (1962). The teachings of these references may be applied to the aforementioned preparation of monobromoperfluoroketones by the reaction of monobromoperfluoroacyl fluorides with perfluoroolefins as well as the aforementioned preparation of monobromoperfluoroketones by the reaction of perfluoroacyl fluorides with monobromoperfluoroolefins. The teachings of these references may also be

applied to the preparation of (perfluoroalkoxy)monobromoperfluoroketones by the reaction of monobromoperfluoroacyl fluorides with perfluorovinyl ethers or by the reaction of perfluoroalkoxyperfluoroacyl fluorides with monobromoolefins. The teachings of these references may also be applied to the preparation of (fluoroalkoxy)monobromoperfluoroketones by the reaction of monobromoperfluoroacyl fluorides with hydrofluorovinyl ethers.

Though not essential for preparing the ketones of the present invention, reaction of a fluoroacyl fluoride (such as a perfluoroacyl fluoride or monobromoperfluoroacyl fluoride) with a fluoroolefin (such as a perfluoroolefin, monobromoperfluoroolefin, perfluorovinyl ether or hydrofluorovinyl ether) may be performed in a polar non-protic solvent such as N,N-dimethyl formamide, N,N-dimethyl acetamide, dimethyl sulfolane, dimethylsulfoxide, N-methylpyrrolidinone, and glycol ethers such as ethylene glycol dimethyl ether, diethylene glycol dimethyl ether, triethylene glycol dimethyl ether, and tetraethylene glycol dimethyl ether. Preferred solvents for reacting fluoroacyl fluorides with fluoroolefin are glycol ethers. The reaction may be run under substantially anhydrous conditions.

The mole ratio of the fluoroolefin to fluoroacyl fluoride during the reaction may be at least about 1:1 to about 2:1, and preferably is about 1.1. More than about 2 moles of fluoroolefin per mole of fluoroacyl fluoride provides little benefit.

The reaction of fluoroacyl fluoride with fluoroolefin is preferably conducted in the presence of a fluoride ion source such as an alkali metal fluoride, alkali metal hydrogen difluoride (i.e., a bifluoride), alkali-earth metal fluoride, tetraalkylammonium fluoride, tetraalkylammonium hydrogen fluoride, trialkylammonium fluoride, or non-oxidizing transition metal fluorides. Preferred fluoride ion sources are potassium fluoride, cesium fluoride, and potassium bifluoride. The fluoride ion source may be present at a level of 5 mole percent to 20 mole percent, preferably about 10 mole percent, based on the quantity of fluoroolefin present.

Temperatures of from about 50° C. to about 250° C., preferably from about 100° C. to about 150° C. are effective to produce any of the fluorinated ketones of the present invention by reaction of a fluoroacyl fluoride with a fluoroolefin.

The reaction of fluoroacyl fluoride with fluoroolefin may take place in batch mode or in semi-batch mode with the fluoroacyl fluoride added gradually to the mixture of the fluoroolefin, solvent, and fluoride ion source. Contact times suitable for the reaction may be from about 0.5 hour to about 24 hours. The reaction typically takes place under autogenous pressure provided by the reactants at the reaction temperature.

Though not added intentionally to the reactions, hydrogen fluoride may be present in small amounts during the reactions of fluoroacyl fluorides due to the presence of traces of water. Reaction of fluoroacyl fluorides with fluoroolefins may be conducted in a vessel formed of materials compatible with hydrogen fluoride at elevated temperatures and pressures. Examples of such materials include stainless steels, in particular of the austenitic type, the well-known high nickel alloys, such as Monel™ nickel-copper alloys, Hastelloy™ nickel-based alloys and, Inconel™ nickel-chromium alloys, and copper-clad steel.

The fluoroketone products may be isolated from the reaction mixture as a lower liquid layer or by distillation. After removing traces of fluoride salts by washing with water, such products may be purified by distillation.

The method of the present invention further includes use of monohydrimonobromoperfluoroketones in which one of the C—F bonds in a perfluoroketone has been replaced by a C—Br bond, and in addition, another of the C—F bonds in said perfluoroketone has been replaced by a C—H bond. New monohydrimonobromoperfluoroketones useful in the process of the present invention comprise $\text{CHF}_2\text{CF}_2\text{C}(\text{O})\text{CBrFCF}_3$, $(\text{CF}_3)_2\text{CHC}(\text{O})\text{CBrFCF}_3$, $\text{CHF}_2\text{CF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$, $(\text{CF}_3)_2\text{CHC}(\text{O})\text{CBr}(\text{CF}_3)_2$, and $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CH}(\text{CF}_3)_2$.

$\text{CHF}_2\text{CF}_2\text{C}(\text{O})\text{CBrFCF}_3$ may be prepared by reacting $\text{CHF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CBrF}=\text{CF}_2$; $(\text{CF}_3)_2\text{CHC}(\text{O})\text{CBrFCF}_3$ may be prepared by reacting $(\text{CF}_3)_2\text{CHC}(\text{O})\text{F}$ with $\text{CBrF}=\text{CF}_2$; $\text{CHF}_2\text{CF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$ may be prepared by reacting $\text{CHF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; $(\text{CF}_3)_2\text{CHC}(\text{O})\text{CBr}(\text{CF}_3)_2$ may be prepared by reacting $(\text{CF}_3)_2\text{CHC}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; and $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CH}(\text{CF}_3)_2$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CH}=\text{CF}_2$. The monohydrimonobromoperfluoroketone $(\text{CF}_3)_2\text{CHC}(\text{O})\text{CBrF}_2$ may be prepared by the reaction of the bromofluoroacyl fluoride $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with the monohydrimonoperfluoroolefin $\text{CF}_3\text{CH}=\text{CF}_2$.

The production of monohydrimonobromoperfluoroketones by the reaction of monohydrimonoperfluoroacyl fluorides with monobromoperfluoroolefins, as well as by the reaction of monobromoperfluoroacyl fluorides with monohydrimonoperfluoroolefins, may use reaction conditions and procedure similar to those discussed hereinabove for the reaction of a fluoroacyl fluoride with a fluoroolefin.

The method of the present invention further includes use of monochloromonobromoperfluoroketones in which one of the C—F bonds in a perfluoroketone has been replaced by a C—Br bond, and in addition, another one of the C—F bonds in said perfluoroketone has been replaced by a C—Cl bond. Monochloromonobromoperfluoroketones useful in the process of the present invention comprise compounds of the formula $\text{CXF}_2\text{CFYC}(\text{O})\text{CFRCF}_3$, wherein X is Cl and Y is Br, or wherein X is Br and Y is Cl, and wherein R is F, a CF_3 radical, or a C_2F_5 radical. These compounds may be prepared by contacting an acid fluoride of the formula $\text{CXF}_2\text{CFYC}(\text{O})\text{F}$, prepared as disclosed by Darst, et al. in U.S. Pat. No. 5,557,010, with a perfluoroolefin of the formula $\text{CFR}=\text{CFR}$.

Representative monochloromonobromoperfluoroketones include $\text{CClF}_2\text{CFBrC}(\text{O})\text{CF}_2\text{CF}_3$, prepared by reacting $\text{CClF}_2\text{CFBrC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CF}_2$; $\text{CBrF}_2\text{CClFC}(\text{O})\text{CF}_2\text{CF}_3$, prepared by reacting $\text{CBrF}_2\text{CClFC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CF}_2$; $\text{CClF}_2\text{CFBrC}(\text{O})\text{CF}(\text{CF}_3)_2$, prepared by reacting $\text{CClF}_2\text{CFBrC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFCF}_3$; and $\text{CBrF}_2\text{CClFC}(\text{O})\text{CF}(\text{CF}_3)_2$, prepared by reacting $\text{CBrF}_2\text{CClFC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CFCF}_3$.

Monochloromonobromoperfluoroketones useful in the method of the present invention further comprise $\text{CClF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$, $\text{CClF}_2\text{CF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$, $\text{CF}_3\text{CClFC}(\text{O})\text{CBr}(\text{CF}_3)_2$, $\text{CClF}_2\text{C}(\text{O})\text{CBrFCF}_3$, $\text{CClF}_2\text{CF}_2\text{C}(\text{O})\text{CBrFCF}_3$, and $\text{CF}_3\text{CClFC}(\text{O})\text{CBrFCF}_3$ which may be prepared by reacting a monochloroperfluoroacyl fluoride with a monobromoperfluoroolefin.

$\text{CClF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$ may be prepared by reacting $\text{CClF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; $\text{CClF}_2\text{CF}_2\text{C}(\text{O})\text{CBr}(\text{CF}_3)_2$ may be prepared by reacting $\text{CClF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; $\text{CF}_3\text{CClFC}(\text{O})\text{CBr}(\text{CF}_3)_2$ may be prepared by reacting $\text{CF}_3\text{CClFC}(\text{O})\text{F}$ with $\text{CF}_3\text{CBr}=\text{CF}_2$; $\text{CClF}_2\text{C}(\text{O})\text{CBrFCF}_3$ may be prepared by reacting $\text{CClF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CBrF}$; $\text{CClF}_2\text{CF}_2\text{C}(\text{O})\text{CBrFCF}_3$ may be prepared by

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reacting $\text{CClF}_2\text{CF}_2(\text{O})\text{F}$ with $\text{CF}_2=\text{CBrF}$; $\text{CF}_3\text{CClFC}(\text{O})\text{CBrFCF}_3$ may be prepared by reacting $\text{CF}_3\text{CClFC}(\text{O})\text{F}$ with $\text{CF}_2=\text{CBrF}$.

Monochloromonobromoperfluoroketones useful in the method of the present invention further include $\text{CBrF}_2\text{C}(\text{O})\text{CCl}(\text{CF}_3)_2$, $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CCl}(\text{CF}_3)_2$, $\text{CBrF}_2\text{C}(\text{O})\text{CClFCF}_3$, and $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CClFCF}_3$ which may be prepared by reacting a monobromoperfluoroacyl fluoride with a monochloroperfluoroolefin.

$\text{CBrF}_2\text{C}(\text{O})\text{CCl}(\text{CF}_3)_2$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CCl}=\text{CF}_2$; $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CCl}(\text{CF}_3)_2$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_3\text{CCl}=\text{CF}_2$; $\text{CBrF}_2\text{C}(\text{O})\text{CClFCF}_3$ may be prepared by reacting $\text{CBrF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CClF}$; and $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{CClFCF}_3$ may be prepared by reacting $\text{CBrF}_2\text{CF}_2\text{C}(\text{O})\text{F}$ with $\text{CF}_2=\text{CClF}$.

The formation of monohydrimonobromoperfluoroketones by the reaction of fluoroacyl fluorides of the formula $\text{CXF}_2\text{CFYC}(\text{O})\text{F}$ with perfluoroolefins, or by the reaction of monochloroperfluoroacyl fluorides with monobromoperfluoroolefins, or by the reaction of monobromoperfluoroacyl fluorides with monochloroperfluoroolefins, may use reaction conditions and procedure similar to those discussed hereinabove for the reaction of a fluoroacyl fluoride with a fluoroolefin.

EXAMPLES

Example 1

Syntheses of $\text{CF}_3\text{CBrFC}(\text{O})\text{C}_2\text{F}_5$ and $\text{CF}_3\text{C}(\text{O})\text{CBrFC}_2\text{F}_5$

Preparation of Perfluoro-2,3-epoxypentane

A 2 L creased flask was equipped with a mechanical stirrer, a thermocouple well, an addition funnel, and a dry ice distillation head connected to a calcium sulfate drying tube. The flask was charged with 800 mL of sodium hypochlorite solution (10-13% chlorine) and 15.0 g of tetrabutylammonium hydrogen sulfate. The addition funnel was charged with 60.0 g (0.24 mole) of cold perfluoro-2-pentene. The solution was cooled to 20° C. using an ice-water bath and the mixture stirred at 600 rpm. The F-2-pentene was then added to the sodium hypochlorite solution over the course of about one hour while maintaining the temperature of the reaction in the range of 20-22° C. After the addition was complete, the mixture was stirred for an additional hour. The flask was then set for distillation and the epoxide product (32.8 g, CAS Reg. No. [71917-15-2]) was then distilled out of the mixture by raising the pot temperature to about 40° C.

Preparation of a Mixture of Perfluoro-2-bromo-3-Pentanone and Perfluoro-3-bromo-2-pentanone

A one liter flask was equipped with a mechanical stirrer, a thermocouple well, and a dry ice condenser connected to a Drierite™ tube. The flask was charged with 146.3 g (1.42 moles) of sodium bromide, 10 g (0.031 mole) of tetrabutyl ammonium bromide, and 235.8 g (300 mL) of acetonitrile. The mixture was stirred for 15 minutes at room temperature and then cooled to about 6° C. using an ice-water bath. A 32.9 g (0.12 mole) sample of perfluoro-2,3-epoxypentane, prepared as described above, was then added to the flask in one portion. The ice bath was removed and the reaction was stirred rapidly for six hours at room temperature.

The flask was then set for vacuum distillation. The bromo-perfluoropentanone mixture (50.3 g) was then distilled out of the flask at a pressure of about 80 mm Hg and a pot temperature about 20-25° C. Analysis of the distillate

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by gas chromatography-mass spectroscopy indicated it was primarily an azeotropic mixture of acetonitrile (31.1 GC area %), perfluoro-2-bromo-3-pentanone (13.6%), and perfluoro-3-bromo-2-pentanone (47.0%).

The distillate was subjected to flash chromatography on silica gel using octane as an eluant. The column effluent containing primarily octane and the mixture of perfluoro-2-bromo-3-pentanone and perfluoro-3-bromo-2-pentanone was then subjected to two vacuum distillations at a pressure of 56 and 100 torr, respectively, to separate the bromoperfluoroketones from the bulk of the octane. The low-boiling fractions from the vacuum distillations were then re-distilled at atmospheric pressure. The fractions collected at a head temperature of 71.2-72.8° C. were combined. The product was a mixture of perfluoro-2-bromo-3-pentanone and perfluoro-3-bromo-2-pentanone in molar ratio of 1.0 to 1.12 with overall purity of >98 GC area %.

Example 2

Fire Extinguishing Concentration of a Mixture of $\text{CF}_3\text{CBrFC}(\text{O})\text{C}_2\text{F}_5$ and $\text{CF}_3\text{C}(\text{O})\text{CBrFC}_2\text{F}_5$

The fire extinguishing concentration of a mixture of $\text{CF}_3\text{CBrFC}(\text{O})\text{C}_2\text{F}_5$ and $\text{F}_3\text{C}(\text{O})\text{CBrFC}_2\text{F}_5$, in a 1.0 to 1.12 mole ratio respectively, was determined by the ICI Cup Burner method. This method is described in "Measurement of Flame-Extinguishing Concentrations" R. Hirst and K. Booth, Fire Technology, vol. 13(4): 296-315 (1977).

Specifically, an air stream is passed at 40 liters/minute through an outer chimney (8.5 cm. I. D. by 53 cm. tall) from a glass bead distributor at its base. A fuel cup burner (3.1 cm. O.D. and 2.15 cm. I.D.) is positioned within the chimney at 30.5 cm. below the top edge of the chimney. The fire extinguishing agent is added to the air stream prior to its entry into the glass bead distributor while the air flow rate is maintained at 40 liters/minute for all tests. The air and agent flow rates are measured using calibrated rotameters.

The test is conducted by adjusting the fuel (n-heptane) level in the reservoir to bring the liquid fuel level in the cup burner just even with the ground glass lip on the burner cup. With the air flow rate maintained at 40 liters/minute, the fuel in the cup burner is ignited. The fire extinguishing agent is added in measured increments until the flame is extinguished.

The fire extinguishing concentration is determined from the following equation: Extinguishing concentration = $(F_1 / (F_1 + F_2)) \times 100$, where F_1 is the agent flow rate and F_2 is the air flow rate.

TABLE 1

FIRE EXTINGUISHING AGENT	FIRE EXTINGUISHING CONCENTRATION (volume % in air)
<u>EXAMPLE</u>	
$\text{CF}_3\text{CBrFC}(\text{O})\text{C}_2\text{F}_5$ and $\text{F}_3\text{C}(\text{O})\text{CBrFC}_2\text{F}_5$ in a 1.0 to 1.12 mole ratio mixture, respectively	3.5
<u>COMPARATIVE</u>	
$\text{CF}_3\text{CHF}_2\text{CF}_3$ (HFC-227ea)	7.3
$\text{CF}_3\text{CHF}_2\text{CHF}_2$ (HFC-236ea)	10.2
$\text{CF}_3\text{CF}_2\text{CH}_2\text{Cl}$ (HCFC-235cb)	6.2
CF_4	20.5
C_2F_6	8.7

TABLE 1-continued

FIRE EXTINGUISHING AGENT	FIRE EXTINGUISHING CONCENTRATION (volume % in air)
CF ₃ Br (Halon-1301)	4.2
CF ₂ ClBr (Halon 1211)	6.2
CHF ₂ Cl	13.6

What is claimed is:

1. A method of extinguishing a fire comprising applying to said fire a composition at least one compound selected from the group consisting of (perfluoroalkoxy)monobromoperfluoroketones and (fluoroalkoxy)monobromoperfluoroketones in an amount sufficient to extinguish said fire.

2. The method of claim 1 wherein said composition comprises a (perfluoroalkoxy)monobromoperfluoroketones.

3. The method of claim 2 wherein said (perfluoroalkoxy)monobromoperfluoroketones is selected from the compounds having the formula R¹C(O)CF(CF₃)OR^F, wherein R¹ is a C₁ to C₃ monobromoperfluoroalkyl radical, and R^F is a C₁ to C₃ perfluoroalkyl radical.

4. The method of claim 3 wherein said (perfluoroalkoxy)monobromoperfluoroketones is selected from the group consisting of: CBrF₂C(O)CF(CF₃)OCF₃, CBrF₂CF₂C(O)CF(CF₃)OCF₃, CBrF₂CF₂CF₂C(O)CF(CF₃)OCF₃, CBrF₂C(O)CF(CF₃)OC₂F₅, CBrF₂CF₂C(O)CF(CF₃)OC₂F₅, CBrF₂C(O)CF(CF₃)OCF₂C₂F₅, CBrF₂CF₂C(O)CF(CF₃)OCF₂C₂F₅, CBrF₂C(O)CF(CF₃)OCF(CF₃)₂, CBrF₂CF₂C(O)CF(CF₃)OCF(CF₃)₂, CF₃CBrFC(O)CF(CF₃)OCF(CF₃)₂, CF₃CBrFC(O)CF(CF₃)OCF₃, (CF₃)₂CBrC(O)CF(CF₃)OCF₃, CF₃CBrFC(O)CF(CF₃)OC₂F₅, (CF₃)₂CBrC(O)CF(CF₃)OC₂F₅, CF₃CBrFC(O)CF(CF₃)OCF₂C₂F₅, and CF₃CBrFC(O)CF(CF₃)OCF(CF₃)₂.

5. The method of claim 1 wherein said composition comprises a (fluoroalkoxy)monobromoperfluoroketones.

6. The method of claim 5 wherein said (fluoroalkoxy)monobromoperfluoroketone is selected from compounds having the formula R¹C(O)CX(CF₃)OR², wherein X is H or F, R¹ is a C₁, C₂ or C₃ bromoperfluoroalkyl radical, and R² is a C₁ to C₃ alkyl or fluoroalkyl radical.

7. The method of claim 6 wherein said (fluoroalkoxy)monobromoperfluoroketone is selected from the group consisting of: CBrF₂C(O)CF(OCF₂CHF₂)CF₃, CBrF₂C(O)

CH(OCF₂CHF₂)CF₃, CBrF₂C(O)CF(OCF₂CHF₂)CF₃, CBrF₂C(O)CH(OCF₂CHF₂)CF₃, CBrF₂C(O)CF(OCH₃)CF₃, and CBrF₂C(O)CF(CF₂OCH₃)CF₃.

8. A method of preventing fire in an air-containing enclosed area containing combustible materials comprising introducing into said area a composition comprising at least one compound selected from the group consisting of (perfluoroalkoxy)monobromoperfluoroketones and (fluoroalkoxy)monobromoperfluoroketones, and maintaining said composition in an amount sufficient to suppress combustion of combustible materials in the enclosed area.

9. The method of claim 8 wherein said composition comprises a (perfluoroalkoxy)monobromoperfluoroketone.

10. The method of claim 9 wherein said (perfluoroalkoxy)monobromoperfluoroketone is selected from the compounds having the formula R¹C(O)CF(CF₃)OR^F, wherein R¹ is a C₁ to C₃ monobromoperfluoroalkyl radical, and R^F is a C₁ to C₃ perfluoroalkyl radical.

11. The method of claim 10 wherein said (perfluoroalkoxy)monobromoperfluoroketone is selected from the group consisting of: CBrF₂C(O)CF(CF₃)OCF₃, CBrF₂CF₂C(O)CF(CF₃)OCF₃, CBrF₂CF₂CF₂C(O)CF(CF₃)OCF₃, CBrF₂C(O)CF(CF₃)OC₂F₅, CBrF₂CF₂C(O)CF(CF₃)OC₂F₅, CBrF₂C(O)CF(CF₃)OCF₂C₂F₅, CBrF₂CF₂C(O)CF(CF₃)OCF₂C₂F₅, CBrF₂C(O)CF(CF₃)OCF(CF₃)₂, CBrF₂CF₂C(O)CF(CF₃)OCF(CF₃)₂, CF₃CBrFC(O)CF(CF₃)OCF(CF₃)₂, CF₃CBrFC(O)CF(CF₃)OCF₃, (CF₃)₂CBrC(O)CF(CF₃)OCF₃, CF₃CBrFC(O)CF(CF₃)OC₂F₅, (CF₃)₂CBrC(O)CF(CF₃)OC₂F₅, CF₃CBrFC(O)CF(CF₃)OCF₂C₂F₅, and CF₃CBrFC(O)CF(CF₃)OCF(CF₃)₂.

12. The method of claim 8 wherein said composition comprises a (fluoroalkoxy)monobromoperfluoroketone.

13. The method of claim 12 wherein said (fluoroalkoxy)monobromoperfluoroketone is selected from compounds having the formula R¹C(O)CX(CF₃)OR², wherein X is H or F, R¹ is a C₁, C₂ or C₃ bromoperfluoroalkyl radical, and R² is a C₁ to C₃ alkyl or fluoroalkyl radical.

14. The method of claim 13 wherein said (fluoroalkoxy)monobromoperfluoroketone is selected from the group consisting of: CBrF₂C(O)CF(OCF₂CHF₂)CF₃, CBrF₂C(O)CH(OCF₂CHF₂)CF₃, CBrF₂C(O)CF(OCH₃)CF₃, and CBrF₂C(O)CF(CF₂OCH₃)CF₃.

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