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(54) **Titre : PROCÉDE DE PRODUCTION DE COMPOSÉS ORGANIQUES FLUORÉS**
(54) **Title: METHOD FOR PRODUCING FLUORINATED ORGANIC COMPOUNDS**

(57) **Abrégé/Abstract:**

Disclosed is a method for producing fluorinated organic compounds, including hydrofluoropropenes, which preferably comprises converting at least one compound of formula (I): $\text{CF}_3\text{CF}_n\text{CH}_m\text{X}_{a-m}$ (I) to at least one compound of formula (II) CF_3CZCHZ (II). where each X is independently Cl, F, I or Br; each Z is independently H or F; n is 1 or 2; m is 1, 2 or 3, provided that when n is 1, m is 1 or 2; a is 2 or 3, and $a-m \geq 0$. Certain embodiments include the step of reacting fluorinated C2 olefin, such as tetrafluoroethylene, with a C1 addition agent under conditions effective to produce a compound of formula (I).



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(54) **Title**: METHOD FOR PRODUCING FLUORINATED ORGANIC COMPOUNDS

(57) **Abstract**: Disclosed is a method for producing fluorinated organic compounds, including hydrofluoropropenes, which preferably comprises converting at least one compound of formula (I): $\text{CF}_3\text{CF}_n\text{CH}_m\text{X}_{a-m}$ (I) to at least one compound of formula (II) CF_3CZCHZ (II). where each X is independently Cl, F, I or Br; each Z is independently H or F; n is 1 or 2; m is 1, 2 or 3, provided that when n is 1, m is 1 or 2; a is 2 or 3, and $a-m \geq 0$. Certain embodiments include the step of reacting fluorinated C2 olefin, such as tetrafluoroethylene, with a Cl addition agent under conditions effective to produce a compound of formula (I).

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METHOD FOR PRODUCING FLUORINATED ORGANIC COMPOUNDS

BACKGROUND OF INVENTION

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(1) Field of Invention:

This invention relates to novel methods for preparing fluorinated organic compounds, and more particularly to methods of producing fluorinated olefins.

(2) Description of Related Art:

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Hydrofluorocarbons (HFC's), in particular hydrofluoroalkenes such tetrafluoropropenes (including 2,3,3,3-tetrafluoro-1-propene (HFO-1234yf) and 1,3,3,3-tetrafluoro-1-propene (HFO-1234ze)) (HFO is hydrofluorolefin) have been disclosed to be effective refrigerants, fire extinguishants, heat transfer media, propellants, foaming agents, blowing agents, gaseous dielectrics, sterilant carriers, polymerization media, particulate removal fluids, carrier fluids, buffing abrasive agents, displacement drying agents and power cycle working fluids. Unlike chlorofluorocarbons (CFCs) and hydrochlorofluorocarbons (HCFCs), both of which potentially damage the Earth's ozone layer, HFCs do not contain chlorine and thus pose no threat to the ozone layer.

20

Several methods of preparing hydrofluoroalkanes are known. For example, U.S. Pat. No. 4,900,874 (Ihara et al) describes a method of making fluorine containing olefins by contacting hydrogen gas with fluorinated alcohols. Although this appears to be a relatively high-yield process, for commercial scale production the handling of hydrogen gas at high temperature raises difficult safety

related questions. Also, the cost of producing hydrogen gas, such as building an on-site hydrogen plant, can be in many situations prohibitive.

U.S. Pat. No. 2,931,840 (Marquis) describes a method of making fluorine containing olefins by pyrolysis of methyl chloride and tetrafluoroethylene or chlorodifluoromethane. This process is a relatively low yield process and a very large percentage of the organic starting material is converted in this process to unwanted and/or unimportant byproducts, including a sizeable amount of carbon black. The carbon black is not only unwanted, it tends to deactivate the catalyst used in the process.

The preparation of HFO-1234yf from trifluoroacetylacetone and sulfur tetrafluoride has been described. See Banks, et al., *Journal of Fluorine Chemistry*, Vol. 82, Iss. 2, p. 171-174 (1997). Also, U.S. Pat. No. 5,162,594 (Krespan) discloses a process wherein tetrafluoroethylene is reacted with another fluorinated ethylene in the liquid phase to produce a polyfluoroolefin product.

Catalyzed hydrogen reduction reactions have been disclosed for the preparation of fluorinated C3 hydrocarbons in U.S. Patent No. 5,545,777. The patent describes the reaction as being one in which a compound of formula (1)



is converted by catalyzed hydrogen reduction to a compound of formula (2)



where a, b, c, x, y and z are integers satisfying the following conditions:

$a \geq 0$, $b \geq 1$, $c \geq 2$, $x \geq 1$, $y \geq 1$, $z \geq 0$, $a+b+c=8$, $x=y+z$, $b-y \geq 0$, and $c-z \geq 2$. Since the reactions disclosed in this patent require a reaction product in which $a+b+c=8$ and

that $x=y+z$, it is not possible for the disclosed reaction product to include C3 olefins, which as mentioned above have been found to be desirable for use in many important applications.

Notwithstanding prior teachings applicants appreciate a continuing need for methods of efficiently preparing certain hydrofluorocarbons, particularly tetrafluoropropenes such as HFO-1234yf.

SUMMARY OF THE INVENTION

Applicants have discovered a method for producing fluorinated organic compounds, including hydrofluoropropenes, which preferably comprises converting at least one compound of formula (I):



to at least one compound of formula (II)



where each X is independently Cl, I or Br; each Z is independently H or F; n is 1 or 2; m is 1, 2 or 3, provided that when n is 1, m is 1 or 2; a is 2 or 3, and $a-m \geq 0$. In certain preferred embodiments, each Z is different. Formula II is intended to include all possible isomers.

The preferred converting step of the present invention comprises catalytic reduction of the compound of formula (I). The catalytic reduction step comprises in preferred embodiments introducing said compound of formula (I) to a reaction system under conditions effective to convert, and preferably convert at least about 50%, more preferably at least about 70%, and even more preferably at least about

90%, of said compound of formula (I). It is also generally preferred that said converting step produces a reaction product having at least about 20% selectivity, more preferably at least about 40% selectivity and even more preferably at least about 70% selectivity, to compounds of formula (II), preferably tetrafluoropropene, and even more preferably HFO-1234yf.

In certain preferred embodiments, the converting step comprises reacting a compound of formula (I) in the gas phase, in the liquid phase, or a combination of these, with gas phase reactions preferably occurring in the presence of catalyst.

DETAILED DESCRIPTION OF PREFERRED EMBODIMENTS

One beneficial aspect of the present invention is that it enables the production of desirable fluoroolefins, preferably C3 fluoroolefins, using relatively high conversion and high selectivity reactions. In addition, the methods of the present invention provided reactions with relatively high yield and which are capable of obtaining relatively long catalyst life.

Furthermore, the present methods in certain preferred embodiments permit the products of the desirable fluoroolefins from relatively attractive starting materials. Ethylene and its halogenated derivatives, such as tetrafluoroethylene, may in certain embodiments be an advantageous starting material because such products are relatively easy to handle, and are generally readily available in commercial quantities and/or can be easily produced from other readily available materials. For example, the compounds of formula (I) can be synthesized by the catalyzed gas phase addition of CH_2FCl and $\text{CF}_2=\text{CF}_2$.

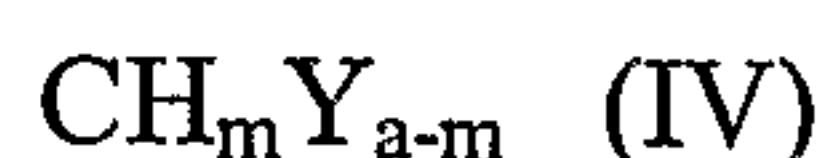
Thus, in certain embodiments the present methods include the step of reacting fluorinated C2 olefin, such as tetrafluoroethylene, with a C1 addition agent under conditions effective to produce a compound of formula (I)



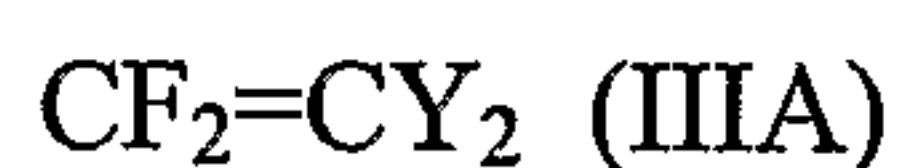
where X, n, m, and a are indicated above. In preferred embodiments, the fluorinated olefin reactant is a compound of formula (III)



where each Y is independently F, Cl, I or Br, provided that at least one F on each carbon atom, and the C1 addition agent comprises a compound of formula (IV)



Where Y is as indicated above and m is 1, 2 or 3; a is 2 or 3, and $a-m \geq 0$. In preferred embodiments, the compound of formula (III) comprises a compound of formula (IIIA)



where each Y is independently F or Cl, and the compound of formula (IV) comprises a compound of formula (IVA)



The reaction by which the compound of formula (III) is converted to a compound of formula (I) is sometimes referred to herein for convenience, but not necessarily by way of limitation, as an addition reaction.

Of course, it is also contemplated that the compounds of formula (I) will, in certain embodiments, themselves be attractive starting materials for the production of the desired fluorinated olefin in accordance with the present

invention, and therefore certain embodiments of the present invention do not include the addition reaction described herein.

Preferably the formula (I) compound, which in certain embodiments is preferably formed by a process comprising a catalyzed C1 addition reaction, is then exposed to reaction conditions effective to produce a reaction product containing one or more of the desired fluorolefins, preferably one or more compounds of formula (II). In one preferred aspect of the present invention, the conversion step comprises a reaction that is sometimes referred to herein for convenience, but not necessarily by way of limitation, as a reduction reaction and in other aspects as a fluorination dehydrohalogenation reaction. Preferred aspects of each of the preferred steps is described below, with the titles used as headings for these steps being used for convenience but not necessarily by way of limitation.

I. ADDITION REACTION

In preferred embodiments, the reactant compound of formula (III) is fluorinated olefin, more preferably fluorinated ethylene and even more preferably $\text{CF}_2=\text{CF}_2$ (sometimes referred to herein as "TFE") or $\text{CF}_2=\text{CFCl}$ (sometimes referred to herein as "CTFE"). Preferably, one or more fluorinated ethylene compounds are reacted with one or more of CH_2FCl .

In certain preferred embodiments, the addition step comprises contacting, (preferably by introducing into a reactor) the compounds in an $\text{CH}_m\text{Y}_{a-m}$:formula III mole ratio of from about 1:1 to about 200:1, more preferably from about 1:1 to

about 100:1 and even more preferably of from about 2:1 to about 3:1. In preferred
embodiments in which the compound of $\text{CH}_m\text{Y}_{a-m}$ comprises CH_2FCl and the
formula III compound comprises $\text{CF}_2=\text{C F}_2$, the CH_2FCl :TFE mole ratio of the
feeds to the reactor are from about 1:1 to about 200:1, more preferably from about
1:1 to about 100:1 and even more preferably from about 1.5:1 to about 2:1.

It is contemplated that this reaction step can be carried out in the liquid
phase or in the gas phase, or a combination of liquid/gas phases, and it is further
contemplated that the reaction can be carried out batch wise, continuous, or a
combination of these.

Thus, it is contemplated that the addition step, when it is used, may be
performed using a wide variety of process parameters and process conditions in
view of the overall teachings contained herein. However, it is preferred in certain
embodiments that this reaction step comprise a gas phase reaction, preferably in
the presence of catalyst, supported on carbon or unsupported, preferably a metal-
based catalyst, such as antimony-based catalysts (such as SbF_3 , SbF_5 , and
partially fluorinated SbCl_3 or SbCl_5) aluminum-based catalyst (such as AlCl_3),
iron-based catalyst such FeCl_3 including such catalysts on a carbon or other
appropriate support. It is expected that many other catalysts may be used
depending on the requirements of particular embodiments, and of course, two or
more any of these catalysts, or other catalysts not named here, may be used in
combination.

The gas phase addition reaction may be conducted, for example, by
introducing a gaseous form of a compound of formula (III) and formula (IV) into

a suitable reaction vessel or reactor. Preferably the vessel is comprised of materials which are resistant to corrosion, such as Hastelloy, Inconel, Monel and/or fluoropolymers linings. Preferably the vessel contains catalyst, for example a fixed or fluid catalyst bed, packed with a suitable addition catalyst, with suitable means to heat the reaction mixture to the desired reaction temperature.

While it is contemplated that a wide variety of reaction temperatures and pressures may be used, depending on relevant factors such as the catalyst being used and the most desired reaction product, it is generally preferred that at least a portion of the addition step is carried out at a reaction temperature of from about 5°C to about 1000°C, more preferably 5°C to about 500°C, more preferably 5°C to about 200°C and even more preferably from about 40°C to about 60°C for reactors which are preferably maintained at a pressure of from about 1 to about 1500 psig, and even more preferably from about 20 to about 40 psig.

In certain preferred embodiments, the compound of formula (III) and the compound of formula (IV) are introduced into an appropriate reaction vessel in the form of a gas and the reactor is preferably maintained at a temperature of about 50°C and the reactor is preferably maintained at a pressure of about 30 psig.

In preferred embodiments the conversion of the formula (III) compound, particularly formula (IIIA) compound(s), is preferably at least about 15%, more preferably at least about 20%, and selectivity to compounds of formula I is preferably at least about 50%, more preferably at least about 70%, and even more

preferably at least about 75%.

II. FORMATION OF THE COMPOUND OF FORMULA II

5 The methods of the present invention preferably comprise converting a compound of formula (I) to a fluorolefin, preferably a C3 fluorolefin, more preferably a compound of formula (II), and even more preferably tetrafluoropropene.

10 In certain preferred embodiments, the present converting step is carried out under conditions effective to provide a formula (I) conversion of at least about 40%, more preferably at least about 55%, and even more preferably at least about 70%. In certain preferred embodiments the conversion is at least about 90%, and more preferably about 100%. Further in certain preferred embodiments, the conversion of the compound of formula I to produce a compound of formula II is
15 conducted under conditions effective to provide a formula II selectivity of at least about 25%, more preferably at least about 40%, more preferably at least about 70%, and even more preferably at least about 90%.

20 This reaction step can be carried out in the liquid phase or in the gas phase, or in a combination of gas and liquid phases, and it is contemplated that the reaction can be carried out batch wise, continuous, or a combination of these.

A. GAS PHASE DEHYDROHALOGENATION

One preferred reaction step in accordance may be described by those reactions in which the compound of formula (I) comprises a compound in which n is 2, that is a compound of formula (IA)



For example, one preferred compound of formula (IA) is 1,1,1,2,2-pentafluoro-3chloropropane ($\text{CF}_3\text{CF}_2\text{CH}_2\text{Cl}$). By way of illustration but not necessarily by way of limitation, an embodiment involving this compound may be shown to proceed by the following reaction, where the reducing agent is hydrogen:



Although applicant does not intend to be bound to or limited by any particular theory of operation, is believed that in some embodiments the above noted reaction proceeds by the formation of another compound in accordance with formula (IA), namely, $\text{CF}_3\text{CF}_2\text{CH}_3$, which is generated as an intermediate or byproduct of the reaction with hydrogen, methane or other dehydrogenating agent. In accordance with such theory, the intermediate formula (IA) compound is then converted to the desired compound of formula (II), preferably HFO-1234yf, under the existing reaction conditions, and preferably on the surface of the catalyst.

In an alternative to the above reaction, the reducing agent comprises methane and the reaction is represented, without limitation, as follows:



In certain preferred embodiments, the stream containing the compound of formula (I), and preferably (IA) is preheated, primarily to avoid condensation, to a temperature of from about 50°C to about 90°C, preferably about 60°C to about 70°C, and introduced into a reaction vessel. The appropriate amount of the reducing agent, which is preferably from about 0.1% to about 500% of the stoichiometric amount, is then added to the reaction vessel. Preferably the vessel is comprised of materials which are resistant to corrosion as Hastelloy, Inconel, Monel and/or fluoropolymers linings. Preferably the vessel contains catalyst, for example a fixed or fluid catalyst bed, packed with a suitable dehydrohalogenation catalyst, with suitable means to heat the reaction mixture to the desired reaction temperature.

Thus, it is contemplated that the dehydrohalogenation reaction step may be performed using a wide variety of process parameters and process conditions in view of the overall teachings contained herein. However, it is preferred in certain embodiments that this reaction step comprise a gas phase reaction, preferably in the presence of catalyst, and even more preferably a carbon- and/or metal-based catalyst, such as activated carbon, palladium on carbon, palladium-based catalyst (including palladium on carbon and palladium on aluminum oxides), and ruthenium-based catalysts (including ruthenium on aluminum oxides). It is expected that many other catalysts may be used depending on the requirements of particular embodiments in view of the teachings contained herein. Of course, two or more any of these catalysts, or other catalysts not named here, may be used in combination.

In general it is preferred that the catalysts are fluorinated, preferably for a period of from about several hours (eg, 6 hours). In preferred embodiments, fluorination of the catalysts comprises exposing the catalyst to a stream of HF at about reaction temperature and under slight pressure, for example about 5-150 psia.

The gas phase dehydrohalogenation reaction may be conducted, for example, by introducing a gaseous form of a compound of formula (I), and preferably (1A) and a gaseous form of the reducing agent (and/or dehydrohalogenation agent), into a suitable reaction vessel or reactor. Preferably the vessel is comprised of materials which are resistant to corrosion as Hastelloy, Inconel, Monel and/or fluoropolymers linings. Preferably the vessel contains catalyst, for example a fixed or fluid catalyst bed, packed with a suitable reduction/dehydrohalogenation catalyst, with suitable means to heat the reaction mixture to the desired reaction temperature.

While it is contemplated that a wide variety of reaction temperatures may be used, depending on relevant factors such as the catalyst being used and the most desired reaction product, it is generally preferred that the reaction temperature for the dehydrohalogenation step, particularly where the formula (I) compound comprises (and even more preferably consists essentially of compounds of formula (1A)) is from about 400°C to about 800°C, preferably about 400°C to about 700°C. For such formula (1A) embodiments in which the reducing agent comprises, and even more preferably consists essentially of hydrogen, the reaction temperature for the dehydrohalogenation step is

preferably from about 450°C to about 600°C, more preferably about from 450°C to about 550°C. For such formula (1A) embodiments in which the reducing agent comprises, and even more preferably consists essentially of methane, the reaction temperature for the dehydrohalogenation step is preferably from about 500°C to about 700°C, more preferably from about 600°C to about 700°C.

In general it is also contemplated that a wide variety of reaction pressures may be used, depending again on relevant factors such as the specific catalyst being used and the most desired reaction product. The reaction pressure can be, for example, superatmospheric, atmospheric or under vacuum, and in certain preferred embodiments is from about 15 to about 120 psia.

In certain embodiments, an inert diluent gas, such as nitrogen, may be used in combination with the other reactor feeds. When such a diluent is used, it is generally preferred that the compound of formula (I) comprise from about 5% to greater than 99% by weight based on the combined weight of diluent and formula (I) compound.

It is contemplated that the amount of catalyst use will vary depending on the particular parameters present in each embodiment. In certain preferred embodiments, the contact time is from about 0.1 seconds to about 1000 second, and preferably from about 2 seconds to about 50 seconds. For embodiments in which the compound of formula (I) comprises or consists essentially of a compound of formula (IA) and the reducing agent comprises or consists essentially of hydrogen, and particularly where the desired product of formula (II) is HFO-1234yf, applicants have found that it is preferred to use as the

catalyst a carbon-based catalyst, such as activated carbon, or palladium-based catalyst, or a catalyst comprising palladium and carbon, such as a palladium on carbon catalyst.

For embodiments in which the compound of formula (I) comprises or consists essentially of a compound of formula (IA) and the reducing agent comprises or consists essentially of methane, and particularly where the desired product of formula (II) is HFO-1234yf, applicants have found that it is preferred to use as the catalyst a carbon-based catalyst, such as activated carbon, or a catalyst based on a Period 6 metal (particularly Cs and Ba) such as BaNO₃ or CsNO₃ (including in combination with aluminum-based catalyst or catalyst support, such as Al₂O₃), or Ni-based catalyst (such as Ni mesh) and combinations of these.

Preferably in such dehydrofluorination embodiments as described in this section, the conversion of the formula (I) compound is at least about 50%, more preferably at least about 65%, and even more preferably at least about 90%. Preferably, the selectivity to HFO-1234yf is at least about 70%, more preferably at least about 80% and more preferably at least about 90%.

B. LIQUID PHASE REDUCTION

One preferred reaction step in accordance may be described by those reactions in which the compound of formula (I) comprises a compound in which n is 1, that is a compound of formula (IB)

CF₃CF=CHX (IB) where X is as previously defined.

For example, one preferred compound of formula (1B) is 3,3,3,2-tetrafluoro-1-chloro-1-propene ($\text{CF}_3\text{CF}=\text{CHCl}$).

It is contemplated that the reduction reaction step may be preformed using a wide variety of process parameters and process conditions in view of the overall teachings contained herein. However, it is preferred in certain embodiments that this reaction step comprise a liquid phase reaction, preferably in the presence of catalyst, and even more preferably in the presence of a catalyst contained in a liquid carrier, such as a solvent for at least one or more of the organic reactants.

Although it is contemplated that many solvents and catalysts will be adaptable for use in connection with these preferred embodiments, it is generally preferred that the solvent comprises tetrahydrofuran, dioxane and the like, and any combinations of solvents including these. In such preferred embodiments, the catalysts preferably include palladium-based catalyst. In certain preferred embodiments the palladium based catalyst comprises, and in certain embodiments consists essentially of tetrakis(triphenylphosphine)palladium(0), $[\text{Pd}(\text{PPh}_3)_4]$ and/or tris(dibenzylideneacetone)dipalladium(0), $\text{Pd}_2(\text{dba})_3$ and combinations of these. In certain preferred embodiments, ligands for the catalyst are included in the reaction mixture, and although many ligands are believed to be adaptable for use with the preferred catalyst systems of the present invention, in certain embodiments the ligands comprise tetra-*n*-butyl phosphine, ammonium formate and combinations of these and/or other ligands. It is expected that other catalysts may be used depending on the requirements of particular embodiments

in view of the teachings contained herein. Of course, two or more any of these catalysts, or other catalysts not named here, may be used in combination.

The liquid phase reduction reaction may be conducted, for example, by introducing the solvent and catalyst into a suitable reaction vessel or reactor.

5 Preferably the vessel is comprised of materials which are resistant to corrosion as Hastelloy, Inconel, Monel and/or fluoropolymers linings.

Prior to its introduction to the reactor, the compound of formula (I), preferably including a compound of formula (IB), is preferably cooled to below its boiling point, and preferably to a temperature of from about -5°C to about 10 20°C and introduced into the solvent. The reaction mixture is then preferably brought to a temperature of from about -20°C to about -50°C, and even more preferably to about -30°C to about -40°C and then a partial vacuum is applied to pull residual air or O₂ from the reactor. The reaction mixture is then preferably heated, preferably with agitation (such as stirring) to a temperature of from about 15 10°C to about 200°C, more preferably from about 20°C to about 150°C, and the reaction mixture is preferably maintained at this temperature for a time period of from about 1 hour to about 48 hours, more preferably from about 15 hours to about 30 hours. During this time period the pressure in the reactor may increase in certain embodiments to about 150 psig to about 250 psig, and even more 20 preferably from about 150 psig to about 200 psig.

It is contemplated that the amount of catalyst use will vary depending on the particular parameters present in each embodiment. In certain preferred embodiments, the weight ratio of the compound of formula (I) to catalyst is from

about 50,000:1 to about 1:1, and even more preferably from about 100:1 to about 1:1.

Preferably in such reduction embodiments as described in this section, the conversion of the formula (IB) compound is at least about 85%, more preferably at least about 95%, and even more preferably about 100%. Preferably, the selectivity to HFO-1234yf is at least about 20%, more preferably at least about 30% and more preferably at least about 40%.

EXAMPLES

Additional features of the present invention are provided in the following examples, which should not be construed as limiting the claims in any way.

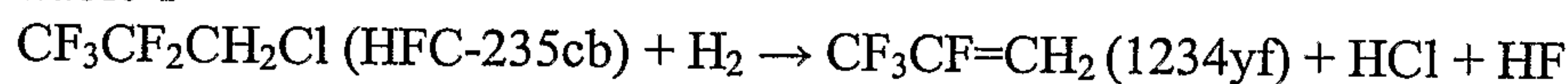
Examples 1 - 11

These examples illustrate gas phase dehydrohalogenation of $\text{CF}_3\text{CF}_2\text{CH}_2\text{Cl}$ (HFC-235cb) to $\text{CF}_3\text{CF}=\text{CH}_2$ (1234yf). A 22-inch (1/2-inch diameter) Monel tube reactor is charged with 120 cc of catalyst, as specified in Table 1 below. The reactor is mounted inside a heater with three zones (top, middle and bottom). The inlet of the reactor is connected to a pre-heater, which was kept at 300°C by electrical heating. Organic (235cb) is fed from a cylinder kept at 65°C. A flow of reducing agent comprising hydrogen gas and the inert N_2 gas is maintained at a rate as indicated in Table 1 below. The reactor temperature is brought to the temperature indicated in the table. The HFC-235cb is passed through gas-flow controllers into a preheater maintained a temperature of about 300°C. The gas stream coming out of the preheater is passed through the catalyst bed at the desired temperature over a specified period of time and at a pressure of

from about 2.5 - 5.3 psig. An on-line GC and a GCMS are used to analyze samples taken at the reactor exit line at regular time intervals. Finally, the reactor effluent is introduced into a 20 - 60 % KOH scrubber solution, and the effluent from the scrubber solution is then condensed to collect the products. The desired product $\text{CF}_3\text{CF}=\text{CH}_2$ (1234yf) is then isolated from the mixture by distillation. The conversion of HFC-235cb is from about 50% to about 100% and the selectivity to HFO-1234yf is from about 60% to about 100%, depending on the reaction conditions. The major byproducts were $\text{CF}_3\text{CF}_2\text{CH}_3$ (HFC-245cb), $\text{CF}_3\text{CF}_2\text{Cl}$ (CFC-115), CF_3Cl (CFC-13), and $\text{CF}_3\text{CF}=\text{CHCl}$.

The results are shown in Table 1 below.

Table 1



Example# / Catalyst	T, °C	H ₂ /N ₂ , Sccm	HFC-235cb, sccm	% Conv. of 235cb	% Selec. to 1234yf	% Selec. to 245cb	% Selec. to R*	% Selec. to R**
Example 1/ A	450	58/0	52	76	60	31	4	4
Example 2/ A	500	61/0	54	100	90	9	0	1
Example 3/ A	550	61/0	54	100	84	6	1	6
Example 4/ A	600	63/0	57	97	63	12	0	24
Example 5/ B	520	41/100	96	80	60	20	5	10
Example 6/ B	550	41/100	138	58	36	40	7	16
Example 7/ C	500	61/0	54	87	69	20	4	6
Example 8/ D	500	61/0	54	80	63	26	3	7

Example 9/ E	500	61/0	63	100	73	16	2	8
Example 10/ F	500	61/0	76	91	30	27	10	30
Example 11/ G	500	61/0	100	80	28	29	8	32

Catalysts (100 cc): A is Calgon activated carbon; B is Shiro-Saga activated carbon; C is Aldrich activated carbon; D is NORIT RFC 3 activated carbon; E is 0.5 wt% Pd/C; F is 0.5 wt% Pd/Al₂O₃; G is 0.5wt% Ru/Al₂O₃

Byproducts: R* is CF₃CF=CHCl and R** is combination of CF₃Cl and CF₃CF₂Cl

Examples 12 - 15

These examples illustrate gas phase hydrodehydrochlorination of CF₃CF₂CH₂Cl to CF₃CF=CH₂ (HFO-1234yf) using methane as the reducing agent. The protocol of Examples 1 – 11 is repeated except as indicated in Table 2 below.

Table 2



Example# / Catalyst	T, °C	P, psig	CH ₄ , sccm	HFC- 235cb, sccm	% Conv. of 235cb	% Selec. to 1234yf
Example 12/ H	560	10	50	50	17	0
Example 13/ I	650	8	100	86	46	26
Example 14/ J	650	7	97	83	52	31

Example# / Catalyst	T, °C	P, psig	CH ₄ , sccm	HFC- 235cb, sccm	% Conv. of 235cb	% Selec. to 1234yf
Example 14*/ K	650 and 500 *	8	96	87	54	40
Example 15/ L	585	3	79	100	77	33

Catalysts (100 cc): H is activated carbon; I 3wt%

BaNO₃/Al₂O₃; J is 3wt% BaNO₃/Al₂O₃ with 0.5wt% CsNO₃ used as a promoter; K is a

two zone reaction with two catalysts, specifically 50 cc CsNO₃ promoted with

BaNO₃/Al₂O₃ is used at 650°C in the first zone and 50 cc Calgon activated carbon is used at

500°C in the second zone; L is Ni mesh.

* two zone reaction

Example 16

This example illustrates the liquid phase reduction of CF₃CF=CHCl to CF₃CF=CH₂ (1234yf) with Pd(PPh₃)₄ catalyst. A 250 mL Parr reactor/autoclave is charged with 30 mL tetrahydrofuran to which 0.5 grams (0.04 mmol) [tetrakis(triphenylphosphine)palladium(0), [Pd(PPh₃)₄], 10.0 grams of ammonium formate (158 mmol), and 11.4 grams cold (0-10 °C) CF₃CF=CHCl (77 mmol) are added under nitrogen. The reactor is sealed immediately, cooled to -30 to -40 °C, and partially evacuated. The contents in the Parr reactor are brought to room temperature and gradually heated to 100°C with stirring. The reactants are maintained at this temperature for 24 hours. During this time, the pressure in the reactor is increased to approximately 180-200 psig. The reactor is then cooled to 25 °C and the volatile materials are collected in an evacuated metal cylinder.

Gas chromatographic (GC) analysis of the volatile materials indicated $\text{CF}_3\text{CF}=\text{CH}_2$ as the main product with trace amounts of $\text{CF}_3\text{CF}=\text{CHCl}$ and carbon dioxide. The reaction mixture is cooled to 0°C and filtrate is analyzed by gas chromatograph (GC) which indicates approximately a 40% conversion of $\text{CF}_3\text{CF}=\text{CHCl}$ to $\text{CF}_3\text{CF}=\text{CH}_2$. Identity of $\text{CF}_3\text{CF}=\text{CH}_2$ is confirmed by comparison with a known sample. Purification of the product is accomplished by passing the product through a cold trap to remove CO_2 and unreacted starting material at about -70°C and then distillation.

Example 17

This example illustrates the liquid phase reduction of $\text{CF}_3\text{CF}=\text{CHCl}$ to $\text{CF}_3\text{CF}=\text{CH}_2$ (HFO-1234yf) with $\text{Pd}_2(\text{dba})_3$ catalyst. A 250 mL Parr reactor/autoclave was charged with 30 mL tetrahydrofuran (30 mL) to which 0.456 grams of tris(dibenzylideneacetone)dipalladium(0), $\text{Pd}_2(\text{dba})_3$ (2.0 mmol), 0.8 grains of tributyl phosphine (2.0 mmol), 8.0 grams of ammonium formate (126 mmol), and 11.4 grams of cold ($5-10^\circ\text{C}$) $\text{CF}_3\text{CF}=\text{CHCl}$ (77 mmol) are added under nitrogen. The reactor is sealed immediately, cooled to the range of from about -30°C to about -40°C , and partially evacuated. The contents in the Parr reactor are brought to room temperature and then gradually heated to 100°C with stirring. The reactants are maintained at this temperature for 24 hours during which the internal pressure rises to approximately 200 psig. The reactor is then cooled to 25°C and the volatile materials are passed through a trap at about -70°C to about -78°C and collected in a cold evacuated metal cylinder. Gas chromatographic (GC) analysis of the volatile materials collected indicate

CF₃CF=CH₂ as the main product with trace amounts of CF₃CF=CHCl and carbon dioxide. The reaction mixture was cooled to 0 °C and filtrate was analyzed by GC which indicated approximately a 35% conversion of CF₃CF=CHCl to CF₃CF=CH₂.

Example 18

This example illustrates the liquid phase reduction of CF₃CF=CHCl to CF₃CF=CH₂ (HFO-1234yf) with Pd₂(dba)₃ catalyst. Example 17 is repeated except that dioxane is substituted for tetrahydrofuran as the solvent. The result is essentially the same as Example 17.

Example 19

This example illustrates the liquid phase reduction of CF₃CF=CHCl to CF₃CF=CH₂ (HFO-1234yf) using Pd₂(dba)₃•CHCl₃ catalyst. The reaction is carried out as in Example 17 except that an equivalent amount of catalyst Pd₂(dba)₃ is substituted by tris(dibenzylideneacetone)dipalladium (0) chloroform complex, Pd₂(dba)₃ •CHCl₃. The extent of conversion of CF₃CF=CHCl to CF₃CF=CH₂ was essentially the same as in Example 17.

Example 20

This example illustrates the liquid phase reduction of CF₃CF=CHCl to CF₃CF=CH₂ (HFO-1234yf) with Pd(PPh₃)₄ catalyst in tetrahydrofuran. A 250 mL Parr Reactor/autoclave was charged with 30 mL of tetrahydrofuran to which 0.5 grams of tetrakis(triphenylphosphine)palladium(0), Pd(PPh₃)₄ (0.04

mmol), 10 grams of ammonium formate (158 mmol), and 11.4 grams of
CF₃CF=CHCl (61 mmol) were added under nitrogen. The reactor is sealed
immediately, cooled to -30 to -40 °C, and partially evacuated. The contents in
the Parr reactor were brought to room temperature and gradually heated to
100°C with constant stirring. The reactants are maintained at this temperature
for 24 hours. The reactor was then cooled to 25°C and the volatile materials
were collected in an evacuated metal cylinder. Gas chromatographic (GC)
analysis of the volatile materials indicated that CF₃CF=CH₂ and CF₃CH=CHCl
are present at a ratio of about 57:42. Carbon dioxide was also present. The
reaction mixture was cooled to 0°C and filtered under pressure, the filtrate was
analyzed by GC which indicated about a 40% conversion of CF₃CF=CHCl to
CF₃CH=CH₂.

Examples 21 - 29

These examples illustrate gas phase dehydrohalogenation of CF₃CF₂CH₃
(HFC-245cb) to CF₃CF=CH₂ (HFO-1234yf). A 22-inch (1/2-inch diameter)
Monel tube reactor is charged with 120 cc of catalyst, as specified in Table I
below. The reactor is mounted inside a heater with three zones (top, middle and
bottom). The inlet of the reactor is connected to a pre-heater, which was kept at
300°C by electrical heating. Organic (HFC-245cb) is fed from a cylinder kept at
65°C. A flow of reducing agent comprising hydrogen gas is maintained at a rate
as indicated in Table 3 below. The reactor temperature is brought to the
temperature indicated in the table. The HFC-245cb is passed through gas-flow

controllers into a preheater maintained a temperature of about 300°C. The gas
 stream coming out of the preheater is passed through the catalyst bed at the
 desired temperature over a specified period of time and at a pressure of from
 about 2.5 – 5.3 psig. An on-line GC and a GCMS are used to analyze samples
 taken at the reactor exit line at regular time intervals. Finally, the reactor effluent
 is introduced into a 20 – 60 % KOH scrubber solution, and the effluent from the
 scrubber solution is then condensed to collect the products. The desired product
 $\text{CF}_3\text{CF}=\text{CH}_2$ (HFO-1234yf) is then isolated from the mixture by distillation. The
 conversion of HFC-245cb is from about 30% to about 70% and the selectivity to
 HFO-1234yf is from about 90% to about 100%, depending on the reaction
 conditions.

The results are shown in Table 3 below.

Table 3

$\text{CF}_3\text{CF}_2\text{CH}_3$ (HFC-245cb) \rightarrow $\text{CF}_3\text{CF}=\text{CH}_2$ (1234yf)

Example# / Catalyst	T, °C	H ₂ /N ₂ , sccm	HFC-245cb, sccm	% Conv. of 245cb	% Selec. to 1234yf
Example 21/ M	575	0	65	79	63
Example 22/ N	575	0	68	82	57
Example 23/ O	575	0	73	73	61
Example 24/ P	575	0	68	84	59

Example# / Catalyst	T, °C	H ₂ /N ₂ , sccm	HFC-245cb, sccm	% Conv. of 245cb	% Selec. to 1234yf
Example 25/ P	575	20	68	89	73
Example 26/ Q	550	0	69	92	53
Example 27/ R	550	0	67	93	33
Example 28/ S	550	0	69	73	46

Catalysts (100 cc): M is NORIT RFC 3; N is Shiro-Saga activated carbon; O is Aldrich activated carbon; P is Calgon activated carbon; Q is 0.5 wt% Pd/C; R is 0.5 wt% PVC; S is Ni-mesh

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Example 29

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This example illustrates the addition formation of compounds of formula (I) by the reaction of TFE with CH₂FCI in a gas phase reaction. Into a ½ inch flow reactor (Monel) 50 grams of freshly prepared catalyst (as indicated below) are charged. CF₂=CF₂ (TFE) and CH₂FCI (R31) are passed through a mass flow controller with a desired flow rate (as indicated below) to the preheater from respective cylinders connected with regulators. The preheater was connected to the reactor and always kept 10°C below the reactor temperature. The reactor was uniformly heated to the desired temperature by an external heating element with an automatic control. The exit line from the reactor was connected to an on-line GC and GCMS for analysis. A 15 wt% KOH scrubber solution was used at 50°C to neutralize acids coming out from the reactor. The gas stream coming out of the scrubber solution was then condensed in a cylinder under liquid N₂ and then finally fractionated (distilled) to isolate products. SbF₅/C and AlCl₃/C are used as the catalyst. At 50°C and under 30 psig reactor pressure, when 50 sccm of TFE and 150 sccm of R31 were passed over SbF₅/C to achieve a 26% conversion of

TFE and an 82% selectivity to $\text{CF}_3\text{CF}_2\text{CH}_2\text{Cl}$. When AlCl_3/C is used as the catalyst, a 35% conversion and 78% selectivity to $\text{CF}_3\text{CF}_2\text{CH}_2\text{Cl}$ was obtained.

Example 30

5 This example illustrates the addition formation of compounds of formula (I) by the reaction of CTFE with CH_2FCl in a gas phase reaction. Example 28 is repeated expect that CTFE is used in place of TFE. The major reaction products include $\text{CF}_3\text{CClFCH}_2\text{Cl}$ and $\text{F}_2\text{ClCCF}_2\text{CH}_2\text{Cl}$ at a 21% conversion (CTFE) level.

Example 31

10 This example illustrates the formation of compounds of formula (I) by the reaction of TFE with CH_2FCl in a gas phase reaction. Into a 300 ml autoclave, 0.1 mol C_2F_4 was reacted with 0.2 mol CH_2ClF in the presence of 0.05 mol of AlCl_3 at 20-30° for 3 hr to give 60% yield to $\text{CF}_3\text{CF}_2\text{CH}_2\text{Cl}$ which was then isolated and
15 purified by distillation.

Example 32

This example illustrates the formation of compounds of formula (I) by the reaction of TFE with CH_2FCl in a gas phase reaction. Into a 300 ml autoclave,
20 0.1 mol C_2F_4 was reacted with 0.2 mol CH_2ClF in the presence of 0.05 mol of AlCl_3 at 20-30° for 3 hr to give 60% yield to $\text{CF}_3\text{CF}_2\text{CH}_2\text{Cl}$ which was then isolated and purified by distillation.

Example 33

25 This example illustrates the formation of compounds of formula (I) by the reaction of CTFE with CH_2FCl in a gas phase reaction. Into a 300 ml

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autoclave, 0.1 mol $\text{C F}_2=\text{CFCl}$ was reacted with 0.2 mol CH_2ClF in the presence of 0.05 mol of AlCl_3 at $20-30^\circ$ for 3 hr to give 60% yield to $\text{CF}_3\text{CClFCH}_2\text{Cl}$ and $\text{F}_2\text{ClCF}_2\text{CH}_2\text{Cl}$ which was then isolated and purified by distillation.

5 The scope of the claims should not be limited by the preferred embodiments and the examples, but should be given the broadest interpretation consistent with the description as a whole.

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CLAIMS:

1. A method for the preparation of 2,3,3,3-tetrafluoro-1-propene ($\text{CF}_3\text{CF}=\text{CH}_2$:
HFO-1234yf comprising: a) adding chlorofluoromethane ($\text{CH}_2\text{FC1}$) to a compound
of formula $\text{CF}_2=\text{CY}_2$, wherein said compound is tetrafluoroethylene ($\text{CF}_2=\text{CF}_2$) or
5 CTFE ($\text{CF}_2=\text{CFC1}$); and b) converting the obtained 1-chloro-2,2,3,3,3-
pentafluoropropane ($\text{CF}_3\text{CF}_2\text{CH}_2\text{C1}$; HCFC-235cb) or 1-chloro-2,3,3,3-tetrafluoro-
1-propene ($\text{CF}_3\text{CF}=\text{CHC1}$) to said HFO-1234yf.
2. The method as claimed in claim 1 wherein said step b) comprises converting
10 the obtained 1-chloro-2,2,3,3,3-pentafluoropropene by a dehydrohalogenation
reaction.
3. The method according to claim 1, wherein the conversion is carried out in the
gas phase.
- 15 4. The method as claimed in claim 3, wherein the conversion is carried out in
the presence of a catalyst.
5. The method according to claim 4, wherein the catalyst is a carbon- and/or
20 metal based catalysts, palladium-based catalysts, or ruthenium-based catalysts or
combinations thereof.
6. The method according to claim 5, wherein the catalyst is activated carbon,
palladium on carbon, palladium on aluminium oxides, or ruthenium on aluminium
oxides or combinations thereof.

7. The method according to claim 1, wherein the conversion is carried out in the liquid phase.
8. The method according to claim 7, wherein the conversion is carried out in the presence of a catalyst.
9. The method according to claim 8, wherein the catalyst is a palladium-based catalyst.
10. The method according to claim 1, wherein step a) is carried out in the gas phase.
11. The method according to claim 10 wherein step a) is carried out in the presence of a catalyst.
12. The method according to claim 11, wherein the catalyst is supported or unsupported and is a metal-based catalyst.
13. The method according to claim 12 wherein the metal-based catalyst is antimony-based catalyst, aluminium-based catalysts and iron-based catalysts.
14. The method as claimed in claim 1 wherein said step b) comprises converting the obtained 1-chloro-2,3,3,3-tetrafluoropropene ($\text{CF}_3\text{CF}=\text{CHCl}$) by a reduction.