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Fortsættes ...

(54) Benævnelse: **FREMGANGSMÅDE TIL FREMSTILLING AF FORMEDE PORØSE PRODUKTER AF CARBON BLACK**

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DESCRIPTION

FIELD OF THE INVENTION

[0001] The present invention generally relates to a method of preparing catalyst compositions comprising a shaped porous carbon product.

BACKGROUND OF THE INVENTION

[0002] Carbon is a material that can be deployed as a catalyst support or adsorbent. The most commonly used carbon based supports for chemical catalysis are activated carbons exhibiting high specific surface areas (e.g., over 500 m²/g). Preparing activated carbon requires activating a carbonaceous material such as charcoal, wood, coconut shell or petroleum-sourced carbon black either by a chemical activation, such as contacting with an acid at high temperatures, or by steam activation. Both methods of activation produce high concentrations of micropores and consequently higher surface areas. Depending upon the source of the carbonaceous material, the resultant activated carbons may have a high residual content of inorganic ash and sulfur, and possibly oxygen or nitrogen-containing functional groups at the surface. Activated carbons are thought to possess an optimum support structure for catalytic applications as they enable good dispersion of catalytically active components and effective adsorption and reaction of chemical reagents at the catalyst surface. Patent documents U.S. 3,978,000; U.S. 4,591,578; U.S. 4,399,052; JP H11100524 A; and CN 103107 319 A describe various carbon black containing compositions and methods of preparing these compositions.

[0003] In recent years, there has been a growing interest in using biorenewable materials as a feedstock to replace or supplement crude oil. See, for example, Klass, Biomass for Renewable Energy, Fuels, and Chemicals, Academic Press, 1998. One of the major challenges for converting biorenewable resources such as carbohydrates (e.g., glucose derived from starch, cellulose or sucrose) to current commodity and specialty chemicals is the selective removal of oxygen atoms from the carbohydrate. Approaches are known for converting carbon-oxygen single bonds to carbon-hydrogen bonds. See, for example, U.S. Patent No. 8,669,397, which describes a process for the conversion of glucose to adipic acid via the intermediate glucaric acid. One challenging aspect associated with the catalytic conversions of highly functionalized biorenewably-derived molecules and intermediates is reaching the high levels of catalytic activity, selectivity and stability necessary for commercial applications. With respect to catalytic activity and selectivity, highly functionalized, biorenewably-derived molecules and intermediates derived from carbohydrates (e.g., glucose and glucaric acid) are non-volatile and must therefore be processed as solutions in the liquid phase. When compared to gas phase catalytic processes, liquid phase catalytic processes are known to suffer from lower productivities because liquid to solid (and gas to liquid to solid) diffusion rates are slower than gas to solid diffusion rates.

[0004] Another challenging aspect associated with the catalytic conversion of highly functionalized biorenewably-derived molecules and intermediates is the use of chemically aggressive reaction conditions. For example, U.S. Patent No. 8,669,397 describes catalytic conversion steps performed at elevated temperatures in the presence of polar solvents such as water and acetic acid. Polar solvents are typically required for the dissolution of non-volatile, highly-functionalized molecules such as glucose and glucaric acid, and elevated temperatures are required for productive and affordable catalytic conversion steps for commodity chemical applications. Therefore, a significant challenge associated with the catalytic conversion of highly functionalized biorenewably-derived molecules and intermediates is catalyst stability.

Long term catalyst stability is a necessity for commodity chemical production, meaning that the catalyst must be stable, productive, and selective under reaction conditions for long periods.

[0005] The challenges associated with the development of industrial shaped catalysts, especially in the conversion of biorenewably-derived molecules and intermediates, are a) high productivity and selectivity consistent with an economically viable catalyst at industrial scale, b) mechanical and chemical stability of the shaped catalyst support and c) retention of the catalytically active components by the support and the avoidance of leaching of the catalytically active components into a polar solvent reaction medium. There remains a need for industrially scalable, highly active, selective and stable catalyst supports and catalyst compositions that can satisfy these challenges.

SUMMARY OF THE INVENTION

[0006] Briefly, the present invention is directed to a method for preparing catalyst compositions comprising a shaped porous carbon product. The method comprises mixing water, carbon black, and a water soluble organic binder, in the absence of an organic solvent, to form a carbon black mixture, wherein the binder comprises: (i) a saccharide selected from the group consisting of a monosaccharide, a disaccharide, an oligosaccharide, and any combination thereof and (ii) a polymeric carbohydrate, wherein the water content of the carbon black mixture is from 5 wt.% to 70 wt.%, the carbon black content of the carbon black mixture is at least 35 wt.% on a dry weight basis, and the binder content of the carbon black mixture is at least 10 wt.% on a dry weight basis, and wherein a binder is deemed water soluble if the solubility at 50°C is at least about 1 wt.%;

forming the carbon black mixture to produce a shaped carbon black composite;

heating the shaped carbon black composite at a temperature of from 300°C to 800°C to carbonize the binder to a water insoluble state and to produce a shaped porous carbon product; and

depositing a catalytically active component or precursor thereof comprising a noble metal on the shaped porous carbon product to produce the catalyst composition.

[0007] For example, the catalyst composition may comprise a shaped porous carbon product as a catalyst support and a catalytically active component or precursor thereof at a surface of the support. The catalyst composition may comprise a shaped porous carbon support and a catalytically active component or precursor thereof comprising platinum and gold at a surface of the support. The catalyst composition may comprise a shaped porous carbon support and a catalytically active component or precursor thereof comprising platinum and rhodium at a surface of the support.

BRIEF DESCRIPTION OF THE DRAWINGS

[0008]

FIG. 1 provides a scanning electron microscopy image of the cross-section of a sample of the catalyst extrudate prepared with Monarch 700 carbon black.

FIG. 2 provides a magnified view of one of catalyst extrudate cross-sections of FIG. 1.

FIG. 3 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a raw Monarch 700 carbon black material.

FIG. 4 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a fresh

catalyst extrudate including the Monarch 700 carbon black material.

FIG. 5 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a catalyst extrudate including the Monarch 700 carbon black material following 350 hours of use.

FIG. 6 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a extrudate using Monarch 700 carbon black and a glucose/hydroxyethyl cellulose binder.

FIG. 7 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a extrudate using Sid Richardson SC159 carbon black and a glucose/hydroxyethyl cellulose binder.

FIG. 8 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a extrudate using Sid Richardson SC 159 carbon black and a glucose/hydroxyethyl cellulose binder which has been exposed to oxygen at 300°C for 3 hours.

FIG. 9 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for an extrudate using Asbury 5368 carbon black and a glucose/hydroxyethyl cellulose binder.

FIG. 10 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for an activated carbon extrudate of Süd Chemie G32H-N-75.

FIG. 11 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for an activated carbon extrudate of Donau Supersorbon K4-35.

FIG. 12 presents the pore size distribution for an extrudate using Sid Richardson SC159 carbon black and a glucose/hydroxyethyl cellulose binder measured by mercury porosimetry.

DETAILED DESCRIPTION

[0009] The present invention relates to a method of preparing a catalyst composition comprising shaped porous carbon products.

[0010] The shaped porous carbon products exhibit high mechanical strength and are resistant to crushing and attrition during use. Further, the shaped porous carbon products possess excellent chemical stability to reactive solvents such as acids and other polar solvents even at elevated temperatures. The shaped porous carbon products are highly suited for liquid phase catalytic reactions because they provide for effective mass transfer of compounds having relatively large molecular volumes to and away the surface of the support.

[0011] The shaped porous carbon products can be prepared from inexpensive and readily available materials which advantageously improves process economics. Furthermore, the disclosed methods are suited for preparation of robust, mechanically strong, shaped porous carbon products through the use of water soluble organic binders. These processes avoid the use of organic solvents that require special handling and storage.

[0012] The shaped porous carbon products exhibit a high degree of retention of the catalytically active component(s) of the catalyst compositions, which beneficially avoids or reduces the amount of catalytically active material leached into a liquid phase reaction medium. Further, the catalyst compositions possess a high degree of stability which is necessary for commodity chemical production.

[0013] Surprisingly, it has been found that the shaped porous carbon products exhibit a superior mechanical strength (e.g., mechanical piece crush strength and/or radial piece crush strength), and the use of catalyst compositions comprising the shaped porous carbon products prepared by the method of the present invention provides unexpectedly higher productivity, selectivity and/or yield in certain reactions when compared to similar catalysts compositions with different catalyst support materials.

Shaped Porous Carbon Products and Methods of Preparation

[0014] The shaped porous carbon products can be prepared with carbon black. Carbon black materials include various subtypes including acetylene black, conductive black, channel black, furnace black, lamp black and thermal black. The primary processes for manufacturing carbon black are the furnace and thermal processes. Generally, carbon black is produced through the deposition of solid carbon particles formed in the gas phase by combustion or thermal cracking of petroleum products. Carbon black materials are characterized by particles with diameters in the nanometer range, typically from 5 to 500 nm. These materials also have much lower surface areas, a higher concentration of mesopores, and lower ash and sulfur content when compared to activated carbons. Carbon black materials are deployed commercially for many applications such as fillers, pigments, reinforcement materials and viscosity modifiers. However, due to their very low surface areas, carbon black materials are not typically used as supports for chemical catalysis or adsorbents. Low surface area carbon black materials can be considered non-optimal as support structures for catalytic applications because low surfaces areas are considered detrimental to effective dispersion of catalytically active components leading to poor catalytic activity.

[0015] As noted, activated carbons are thought to possess an optimum support structure for catalytic applications as they enable good dispersion of catalytically active components and effective adsorption and reaction of chemical reagents at the catalyst surface. In contrast, the use of carbon black as a catalyst support has been limited. In order to utilize carbon blacks as supports for chemical catalysis, several groups have reported methods to modify carbon black materials. Reported modifications are centered on methods to increase the surface area of the carbon black materials. U.S. Patent No. 6,337,302 describes a process to render a "virtually useless" carbon black into an activated carbon for commodity applications. U.S. Patent No. 3,329,626 describes a process to convert carbon black materials with surface areas from 40-150 m²/g by steam activation into activated carbons with surface areas up to around 1200 m²/g.

[0016] Notwithstanding these teachings, it has been surprisingly discovered that certain carbon black materials exhibiting particular combinations of characteristics such as surface area, pore volume, and pore diameter are highly effective for use in shaped porous carbon catalyst supports for catalytic reactions including liquid and mixed phase reaction mediums. The shaped porous carbon products can be shaped into mechanically strong, chemically stable robust forms that can reduce resistance to liquid and gas flows, withstand desired process conditions, and provide for long term, stable catalytic operation. These shaped porous carbon products provide high productivity and high selectivity during long term continuous flow operation under demanding reaction conditions including liquid phase reactions in which the catalyst composition is exposed to reactive solvents such as acids and water at elevated temperatures.

[0017] Carbon black may constitute a large portion of the shaped porous carbon product of the present invention. As such, the carbon black content of the shaped porous carbon product is at least 35 wt.%, such as from 35 wt.% to 80 wt.%.

[0018] Typically, the carbon black materials used to prepare a shaped porous carbon product have a BET specific surface area in the range of from 20 m²/g to 500 m²/g. The specific surface area of carbon black

materials is determined from nitrogen adsorption data using the Brunauer, Emmett and Teller (BET) Theory. See J. Am. Chem. Soc. 1938, 60, 309-331 and ASTM Test Methods ASTM 3663, D6556, or D4567 which are Standard Test Methods for Total and External Surface Area Measurements by Nitrogen Adsorption.

[0019] The carbon black materials generally have a mean pore diameter greater than 5 nm, such as from 5 nm to 100 nm. Such pore diameters enable effective transport of reactant molecules possessing large molecular volumes (such as biorenewably-derived molecules with 6-carbon atom frameworks) into and out of the pore structure of the catalytically active surface, thereby enabling enhanced activity.

[0020] The carbon black materials used to prepare the shaped porous carbon products also generally have specific pore volumes greater than 0.1 cm³/g, such as from 0.1 cm³/g to 1 cm³/g. Carbon black materials with these specific pore volumes provide a volume sufficient to provide uniform wetting and good dispersion of the catalytically active components while enabling sufficient contact between the reactant molecules and the catalytically active surface. Mean pore diameters and pore volumes are determined in accordance with the procedures described in E.P. Barrett, L.G. Joyner, P. P. Halenda, J. Am. Chem. Soc. 1951, 73, 373-380 (BJH method), and ASTM D4222-03(2008) Standard Test Method for Determination of Nitrogen Adsorption and Desorption Isotherms of Catalysts and Catalyst Carriers by Static Volumetric Measurements.

[0021] Certain carbon black materials are known to be electrically conductive. Accordingly, the shaped porous carbon product may comprise conductive carbon black and may be electrically conductive. The shaped porous carbon product may alternatively comprise nonconductive carbon black and the shaped porous carbon product may not exhibit a conductivity that is suitable for a conductive electrode. The shaped porous carbon product may comprise nonconductive carbon black and less than 50% conductive carbon black based on the total weight of the carbon black in the shaped porous carbon product and/or the total weight of the carbon black used to prepare the shaped porous carbon product. The shaped porous carbon product may comprise carbon black consisting of or consisting essentially of nonconductive carbon black. The carbon black may comprise a silica-bound or alumina-bound carbon black. The shaped porous carbon product may further include graphite and/or a metal oxide (e.g., alumina, silica, titania, and the like).

[0022] The shaped porous carbon product comprising carbon black may be prepared by various methods such as dry powder pressing, drip casting, injection molding, 3D-printing, extrusion and other pelletizing and granulating methods. For example, dry powder pressing involves compressing carbon black particles in a press such as a hot or cold isostatic press or a calendering press. Other pelletizing and granulating methods include tumbling carbon black particles and contacting the particles with a spray containing a binder.

In the method of the present invention the catalyst is prepared by mixing water, carbon black, and a water soluble organic binder, in the absence of an organic solvent, to form a carbon black mixture, wherein the binder comprises: (i) a saccharide selected from the group consisting of a monosaccharide, a disaccharide, an oligosaccharide, and any combination thereof and (ii) a polymeric carbohydrate, wherein the water content of the carbon black mixture is from 5 wt.% to 70 wt.%, the carbon black content of the carbon black mixture is at least 35 wt.% on a dry weight basis, and the binder content of the carbon black mixture is at least 10 wt.% on a dry weight basis, and wherein a binder is deemed water soluble if the solubility at 50°C is at least about 1 wt.%;

forming the carbon black mixture to produce a shaped carbon black composite;

heating the shaped carbon black composite at a temperature of from 300°C to 800°C to carbonize the binder to a water insoluble state and to produce a shaped porous carbon product; and

depositing a catalytically active component or precursor thereof comprising a noble metal on the shaped porous carbon product to produce the catalyst composition.

[0023] A binder solution can be prepared by mixing water and the binder prior to mixing with carbon black. Typically, the binder solution and carbon black mixture are relatively concentrated in binder. The water content of the carbon black mixture is from 5 wt.% to 70 wt.%. The viscosity of the binder solution can vary, for example, according to the binder content and can be readily adjusted to suit a particular shaping process by varying the relative quantities of solid and liquid components. For example, the viscosity of the aqueous solution can be varied by adjusting the amount of binder and type of binder utilized. Also, the water and binder can be mixed and heated to form the binder solution. In some instances, heating can enhance the amount of binder that can be incorporated into the binder solution and/or carbon black mixture (e.g., by increasing the solubility of the binder). For example, the water and binder can be heated to a temperature of at least 50°C, such as from 50°C to 95°C.

[0024] After mixing and heating to form the binder solution, the binder solution can be cooled as needed prior to mixing with carbon black or prior to forming the shaped carbon black composite.

[0025] One method of preparing the shaped porous carbon product comprises mixing carbon black particles with a solution comprising a binder to produce a slurry; forming the slurry (e.g., by extrusion) to produce a shaped carbon black composite and heating or pyrolyzing the shaped carbon black composite to carbonize the binder to produce the shaped porous carbon product.

[0026] In methods of preparing the shaped porous carbon product as described herein, a binder solution or binder and water are thoroughly mixed and blended with the carbon black to prepare a carbon black mixture (e.g., a slurry or a paste). The weight ratio of binder to carbon black in the carbon black mixture is typically at least 1:4. The weight ratio of binder to carbon black in the carbon black mixture can be from 1:4 to 1:1. The carbon black content of the carbon black mixture is at least 35 wt.%, such as from 35 wt.% to 80 wt.% on a dry weight basis. Also, the binder content of the carbon black mixture is at least 10 wt.%, such as from 10 wt.% to 50 wt.% on a dry weight basis.

[0027] Methods of preparing the shaped porous carbon products can further comprise pressing or kneading the carbon black mixture. Pressing or kneading the carbon black mixture compacts the mixture and can reduce the water content of the mixture. Pressing or kneading of the water, carbon black and binder (carbon black mixture) can be conducted simultaneously with the mixing of the water, carbon black and binder. For example, one method of mixing the water, carbon black, and water soluble binder in the absence of an organic solvent and simultaneously pressing the resulting carbon black mixture can be conducted using a mixer muller.

[0028] After mixing of the carbon black and binder, the resulting carbon black mixture is formed into a shaped carbon black composite structure of the desired shape and dimensions by a forming technique such as extrusion, pelletizing, pilling, tableting, cold or hot isostatic pressing, calendering, injection molding, 3D printing, drip casting, or other methods known to produce shaped structures. Forming methods such as cold or hot isostatic pressing and 3D printing may or may not require a binder.

[0029] In general, the shaped porous carbon product can be shaped and sized for use in known industrial reactor formats such as batch slurry, continuous slurry-based stirred tank reactors, fixed beds, ebulated beds and other known industrial reactor formats. The shaped porous carbon product may be formed into various shapes including spheres, beads, cylinders, pellets, multi-lobed shapes, rings, stars, ripped cylinders, triholes, alphas, wheels, etc. Also, the shaped porous carbon product may be formed into amorphous, non-geometric, and random shapes as well as unsymmetrical shapes like hiflow rings and cones and alpha-rings. The mean diameter of the shaped porous carbon product is typically at least 50 μm (0.05 mm) or larger to accommodate process requirements.

[0030] For extrusion forming, a pressure of at least 100 kPa (1 bar) or between 100 kPa (1 bar) to 10,000 kPa (100 bar) is typically applied to the carbon black mixture.

[0031] In drip casting methods, the carbon black mixture comprising carbon black particles and the binder are dispensed as droplets into a casting bath to form the shaped carbon black composite, which is then separated from the casting bath. Carbon black mixture droplets of a tailored diameter may be dispensed through a sized nozzle and dropped into a bath to produce solidified, spherically-shaped carbon black composite of various diameters. In this method, the binder may comprise an alginate (or alginate in combination with another carbohydrate binder as described herein) which can be dispensed into a bath containing a reagent to cause solidification such as an ionic salt (e.g., calcium salt) as described in U.S. Patent No. 5,472,648. The droplets are subsequently allowed to remain substantially free in the ionic solution until the required degree of solidification and consolidation has been attained. Alternatively, the drip casting bath utilized may be, for example, an oil bath, or a bath to cause freeze drying. When an oil bath is used, the temperature of the oil is sufficiently high that the binder is thermally set (e.g., causes the binder to convert to a three-dimensional gel). When a freeze drying bath is used, the resultant beads are typically dried by vacuum treatment. The shaped carbon black composites resulting from such dip casting methods are subsequently pyrolyzed.

[0032] As described in further detail below, other components can be added to the carbon black mixture to assist with the shaping process (e.g., lubricants, compatibilizers, etc.) or to provide other benefits. The carbon black mixture may further comprise a forming adjuvant. For example, the forming adjuvant can comprise a lubricant. Suitable forming adjuvants include, for instance, lignin or lignin derivatives.

[0033] Further, porogens may be mixed with the carbon black and binder to modify and attain the desired pore characteristics in the shaped porous carbon product. Other methods of modifying the porosity of the shaped porous carbon product include mixing two or more different carbon black starting materials (e.g., carbon blacks having different shape and/or size that pack irregularly resulting in multimodal pore size distributions, or carbon blacks from different sources/suppliers, or mixing carbon black powders carbon). Other methods of modifying the porosity of the shaped porous carbon product include multiple thermal processing and/or multiple compounding (e.g., pyrolysis of a shaped carbon black composite of carbon powder and binder, then mixing with fresh carbon black powder and binder and pyrolyzing the resultant composite again).

[0034] In various methods of preparing the shaped porous carbon product, after processing the carbon black mixture (e.g., a slurry or a paste) into the shaped carbon black composite, the composite may be dried to dehydrate the composite. Drying may be achieved by heating the composite at atmospheric pressure and temperatures typically of from room temperature (e.g., 20°C) to 150°C. Other methods of drying may be utilized including vacuum drying, freeze drying, and desiccation. When using certain preparation methods for forming (e.g., tableting, pressing), no drying step may be required.

[0035] In the method of the invention, the shaped carbon black composite (e.g., resulting from extrusion, pelletizing, pilling, tableting, cold or hot isostatic pressing, calandring, injection molding, 3D printing, drip casting, and other forming methods) is heat treated in an inert (e.g., an inert nitrogen atmosphere), oxidative, or reductive atmosphere to carbonize at least a portion of the binder to a water insoluble state and produce a shaped porous carbon product. The heat treatment is typically conducted at a temperature of from 300°C to 800°C. In some instances, and depending upon the binder employed, it has been determined that lower carbonization temperatures can lead to slow leaching of the remnants of the binder from the shaped porous carbon product which reduces mechanical strength over extended periods of use in catalytic reactions. Generally, to ensure longer term stability, the heat treatment is conducted at higher

carbonization temperatures within the ranges specified above. In some instances, the resultant shaped porous carbon product may be washed after the heat treatment to remove impurities.

[0036] In accordance with various preparation methods, shaped porous carbon products comprise a binder or carbonization product thereof in addition to carbon black. Various references, including U.S. Patent No. 3,978,000, describe the use of acetone soluble organic polymers and thermosetting resin as binders for shaped carbon supports. However, the use of flammable organic solvents and expensive thermosetting resins is not desirable or economical for manufacturing large quantities of shaped porous carbon product.

[0037] Mechanically-strong, shaped porous carbon products of the invention can be prepared by the use of an effective binder. The use of an effective binder provides a robust shaped porous carbon product capable of withstanding the prevailing conditions within continuous liquid phase flow environments such as in conversions of biorenewably-derived molecules or intermediates in which the liquid phase may contain water or acidic media. In such instances the shaped porous carbon product is mechanically and chemically stable to enable long-term operation without significant loss in catalyst performance. Moreover, the use of an effective binder provides a robust shaped porous carbon product capable of withstanding elevated temperatures.

[0038] Applicants have found that readily available water soluble organic compounds are suitable binders for the preparation of mechanically strong shaped porous carbon products. As used in herein, a binder is deemed water soluble if the solubility at 50°C is at least 1 wt.%. Aqueous solutions of organic binders are highly amenable to commercial manufacturing methods. Organic binders that dissolve in aqueous solutions enable good mixing and dispersion when contacted with the carbon black materials. These binders also avoid safety and processing issues associated with large-scale use of organic solvents which may be flammable and require special storage and handling. Also, these binders are relatively inexpensive when compared to costly polymer-based binders. As such, the carbon black mixtures do not contain water immiscible solvents.

[0039] The carbohydrate binder may be used in the form of syrups such as molasses or corn syrups or soluble starches or soluble gum or modified versions thereof.

[0040] The water soluble organic binder used in the method of the present invention comprises (i) a saccharide selected from the group consisting of a monosaccharide, a disaccharide, an oligosaccharide, and any combination thereof and (ii) a polymeric carbohydrate. The weight ratio of (i) the saccharide to (ii) the polymeric carbohydrate, can be from 5:1 to 50:1.

[0041] The water soluble organic binder may comprise a monosaccharide. For example, the monosaccharide can be selected from the group consisting of glucose, fructose, hydrates thereof, syrups thereof (e.g., corn syrups, molasses, and the like) and combinations thereof. The water soluble organic binder may comprise a disaccharide. Disaccharides include for example, maltose, sucrose, syrup thereof, and combinations thereof.

[0042] As noted, the binder comprises a polymeric carbohydrate. The polymeric carbohydrate may comprise a cellulosic compound. Cellulosic compounds include, for example, methylcellulose, ethylcellulose, ethylmethylcellulose, hydroxyethylcellulose, hydroxypropylcellulose, methylhydroxyethylcellulose, ethylhydroxyethylcellulose, hydroxypropylmethylcellulose, carboxymethylcellulose, and mixtures thereof.

[0043] Further, the polymeric carbohydrate may be selected from the group consisting of alginic acid, pectin, aldonic acids, aldaric acids, uronic acids, sugar alcohols, and salts, oligomers, and polymers

thereof. The polymeric carbohydrate may also comprise a starch or a soluble gum.

[0044] The water soluble organic binder may comprise a cellulosic compound. The binder may comprise an acidic polysaccharide such as alginic acid, pectin or a salt thereof or the binder may comprise a soluble cellulose such as an alkyl cellulose (e.g., hydroxyethylcellulose, hydroxypropylcellulose, hydroxyethylmethylcellulose, hydroxypropylmethylcellulose, and carboxymethylcellulose).

[0045] As described above, water soluble organic binders that may be used in combination with the saccharide binders include water soluble celluloses and starches (e.g., hydroxyethylcellulose, hydroxypropylcellulose, hydroxyethylmethylcellulose, hydroxypropylmethylcellulose, carboxymethylcellulose), water soluble alcohols (e.g., sorbitol, xylitol, polyvinylalcohols), water soluble acetals (e.g., polyvinylbutyral), water soluble acids (e.g., stearic acid, citric acid, alginic acid, aldonic acids, aldaric acids, uronic acids, xylonic or xylaric acids (or oligomers, or polymers or salts or esters thereof) polyvinyl acrylic acids (or salts or esters thereof). The combination of water soluble organic binders may comprise a cellulosic compound and a monosaccharide. The cellulosic compound may comprise hydroxyethylcellulose, or methylcellulose and the monosaccharide comprises a glucose, fructose or hydrate thereof (e.g., glucose). In particular, one combination comprises glucose and hydroxyethylcellulose, which provides shaped porous carbon products with enhanced mechanical strength, particularly when processed at high carbonization temperatures. The combination of water soluble organic binders may comprise a monosaccharide and a water-soluble alcohol such as sorbitol, mannitol, xylitol or a polyvinyl alcohol. The combination of water soluble organic binders may comprises a monosaccharide, and a water-soluble acid such as stearic acid, pectin, or alginic acid.

[0046] Other water soluble compounds may be used in combination with a carbohydrate or polymeric binder. Combining a carbohydrate or other binder with selected other water soluble organic compounds can provide advantages in the preparation of and in the properties of the resultant shaped porous carbon product. For example, water soluble organic compounds such as stearic acid or stearates such as Zr or NH_4 stearate can provide lubrication during the forming process. Wetting agents may be added (e.g., GLYDOL series available commercially from Zschimmer and Schwarz).

[0047] Porogens may also be added in combination with the binder (or binders). Porogens are typically added to occupy a specific molecular volume within the formulation such that after the shaping and thermal processing the porogen will be pyrolyzed leaving pores of a certain volume and diameter within the shaped product. The presence of such pores can be beneficial to performance. For example, when used as a catalyst support the presence of such pores can lead to more efficient diffusion (of reactants and products) to and from the catalytically active surfaces. More efficient access and egress for the reactants and products can lead to improvements in catalyst productivity and selectivity. Porogens are typically oligomeric (e.g., dimer, trimers of higher order oligomers) or polymeric in nature. Water soluble organic compounds such as water soluble linear and branched polymers and cross-linked polymers are suitable for use as porogens. Polyacrylates (such as weakly cross-linked polyacrylates known as superabsorbers), polyvinyl alcohols, polyvinylacetates, polyesters, polyethers, or copolymers (which may be block copolymers) thereof may be used as porogens. In some instances, the water soluble copolymer may be a block copolymer comprising a water soluble polymer block and a second polymer block which may be hydrophobic and amenable to carbonization (e.g., polystyrene). In another embodiment, polymer dispersions in water are used as binders, i.e., non-water soluble polymers dispersed in water (with the aid of surfactants) such as commercial polyvinyl alcohol, polyacrylonitrile, polyacrylonitrile-butadiene-styrene, phenolic polymer dispersions. Also, copolymers consisting of a water-soluble branch (e.g., polyacrylic acid) and a hydrophobic branch (e.g., polymaleic anhydride, polystyrene) enabling water solubility of the copolymer and enabling carbonization of the hydrophobic branch without depolymerisation upon pyrolysis. Carbohydrates or derivatives thereof, (disaccharides, oligosaccharides, polysaccharides such as sucrose,

maltose, trihalose, starch, cellubiose, celluloses), water soluble polymers and polymer dispersions in water may be used together in any combination as a porogen to attain a shaped porous carbon black product having the desired pore size and volume characteristics described herein.

[0048] Porogens can also be added as gels (e.g., pre-gelated superabsorber) or water insoluble incompressible solids (e.g., polystyrene microbeads, lignins, phenolic polymers) or expandable porogens such as EXPANSEL microspheres available from Akzo Nobel Pulp and Performance (Sundsvall, Sweden). The molecular weight of the oligomer or polymer can be also chosen to design a desired pore sizes and volume characteristics of the shaped carbon product of the invention. For example, the desired shaped carbon product may have a monomodal, bimodal or multimodal pore size distribution as a consequence of addition of a porogen. For illustration, a bimodal or multimodal pore size distribution may consist of a high percentage of pores between 10 and 100 nm and additionally the presence of pores >100 nm. Such a pore structure may provide performance advantages. For example, the presence of a such a pore size distribution can lead to more efficient diffusion (of reactants and products) through the larger pore (transport pores) to and from catalytically active surfaces which reside in the pores sized between 10 and 100 nm. More efficient access and egress for the reactants and products can lead to improvements in catalyst productivity, selectivity, and/or yield.

[0049] Following heat treatment of the shaped carbon black composite, the resulting shaped porous carbon product comprises carbon black and carbonized binder. More generally, the shaped porous carbon product can comprise a carbon agglomerate. Without being bound by any particular theory, it is believed that the carbon agglomerate comprises carbon aggregates or particles that are physically bound or entangled at least in part by the carbonized binder. Moreover, and without being bound by any particular theory, the resulting agglomerate may include chemical bonding of the carbonized binder with the carbon aggregates or particles.

[0050] The carbonized binder comprises a carbonization product of a water soluble organic binder as described herein. Carbonizing the binder during preparation of the shaped porous carbon product may reduce the weight of the shaped carbon black composite from which it is formed. Accordingly, in various embodiments, the carbonized binder content of the shaped porous carbon product is from 10 wt.% to 50 wt.%. The specific surface area (BET surface area), mean pore diameter, and specific pore volume of the shaped porous carbon products are generally comparable to that exhibited by the carbon black material used to prepare the products. However, the preparation process can lead to a reduction or an increase in these characteristics of the products as compared to the carbon black material (e.g., a 10-50% decrease or increase). In various embodiments, the shaped porous carbon product has a specific surface area from 20 m²/g to 500 m²/g. The specific surface area of the shaped porous carbon product is determined from nitrogen adsorption data using the Brunauer, Emmett and Teller. See the methods described in J. Am. Chem. Soc. 1938, 60, 309-331 and ASTM Test Methods D3663, D6556 or D4567, which are Standard Test Methods for Surface Area Measurements by Nitrogen Adsorption.

[0051] The shaped porous carbon products typically have a mean pore diameter greater than 5 nm, such as from 10 nm to 100 nm. Also, the shaped porous carbon products have specific pore volumes of the pores having a diameter of 1.7 nm to 100 nm as measured by the BJH method that is generally greater than 0.1 cm³/g, such as from 0.1 cm³/g to 1.5 cm³/g. Mean pore diameters and specific pore volumes are determined in accordance with the procedures described in E.P. Barrett, L.G. Joyner, P. P. Halenda, J. Am. Chem. Soc. 1951, 73, 373-380 (BJH method), and ASTM D4222-03(2008) Standard Test Method for Determination of Nitrogen Adsorption and Desorption Isotherms of Catalysts and Catalyst Carriers by Static Volumetric Measurements.

[0052] It has been observed that the magnitude of the specific surface area is generally proportional to the concentration of micropores in the shaped porous carbon product structure. In particular, the shaped porous carbon products generally possess a low concentration of pores having a mean diameter less than 1.7 nm. Typically, pores having a mean diameter less than 1.7 nm constitute no more than 10% of the pore volume of the shaped porous carbon product. Similarly, in various embodiments, the pore size distribution of the shaped porous carbon products is such that peaks below 5 nm are not observed. For example, the shaped porous carbon products can have a pore size distribution such that the peak of the distribution is at a diameter greater than 5 nm. Also, the shaped porous carbon product can have a pore size distribution such that the peak of the distribution is at a diameter less than 100 nm.

[0053] Moreover, the shaped porous carbon product advantageously exhibits a high concentration of mesopores between 10 nm to 100 nm. Accordingly, at least 50% of the pore volume of the shaped porous carbon product, as measured by the BJH method on the basis of pores having a diameter from 1.7 nm to 100 nm, may be attributable to pores having a mean pore diameter of from 10 nm to 100 nm. Also, at least 35% may be attributable to pores having a mean pore diameter of from 10 nm to 50 nm. Typically, the shaped porous carbon product exhibits a relatively low concentration of pores less than 10 nm. For example, no more than 10% of the pore volume of the shaped porous carbon product, as measured by the BJH method on the basis of pores having a diameter from 1.7 nm to 100 nm, may be attributable to pores having a mean pore diameter less than 10 nm. As an example, from 0.1% to 10% of the pore volume of the shaped porous carbon product, as measured by the BJH method on the basis of pores having a diameter from 1.7 nm to 100 nm, may be attributable to pores having a mean pore diameter less than 10 nm.

[0054] The shaped porous carbon products described herein are mechanically strong and stable. Crush strength represents the resistance of a solid to compression, and is an important property in the industrial use of the shaped porous carbon product as described herein. Instruments for measuring the piece crush strength of individual solid particles generally include a dynamometer that measures the force progressively applied to the solid during the advancement of a piston. The applied force increases until the solid breaks and collapses into small pieces and eventually powder. The corresponding value of the collapsing force is defined as piece crush strength and is typically averaged over multiple samples. Standard protocols for measuring crush strength are known in the art. For example, the mechanical strength of the shaped porous carbon product can be measured by piece crush strength test protocols described by ASTM D4179 or ASTM D6175. Some of these test methods are reportedly limited to particles of a defined dimensional range, geometry, or method of manufacture. However, crush strength of irregularly shaped particles and particles of varying dimension and manufacture may nevertheless be adequately measured by these and similar test methods.

[0055] The shaped porous carbon product may have a radial piece crush strength of greater than 4.4 N/mm (1 lb/mm). The radial piece crush strength of the shaped porous carbon product may be from 4.4 N/mm (1 lb/mm) to 88 N/mm (20 lbs/mm). In radial piece crush strength measurements, the measured force is relative to the dimension of the solid perpendicular to the applied load, which typically can range from 0.5 mm to 20 mm. For irregularly shaped porous carbon products, the radial piece crush strength is measured by applying the load perpendicular to the longest dimension of the solid.

[0056] Mechanical piece crush strength can also be reported on a basis that is unitless with respect to the dimension of the shaped porous carbon product (e.g., for generally spherically shaped solids or solids having approximately equal transverse dimensions). The shaped porous carbon product prepared in accordance with the present invention typically has a piece crush strength greater than 22 N (5 lbs), such as from 22 N (5 lbs) to 88 N (20 lbs).

[0057] In addition to crush strength, the shaped porous carbon products also exhibit desirable attrition and

abrasion resistance characteristics. There are several test methods suitable for determining the attrition and abrasion resistance of the shaped porous carbon products and catalysts produced in accordance with the present disclosure. These methods are a measure of the propensity of the material to produce fines in the course of transportation, handling, and use on stream.

[0058] One such method is the attrition index as determined in accordance with ASTM D4058-96 (Standard Test Method for Attrition and Abrasion of Catalysts and Catalyst Carriers), which is a measurement of the resistance of a material (e.g., extrudate or catalyst particle) to attrition wear due to the repeated striking of the particle against hard surfaces within a specified rotating test drum. This test method is generally applicable to tablets, extrudates, spheres, granules, pellets as well as irregularly shaped particles typically having at least one dimension larger than 1/16 in. (1.6 mm) and smaller than 3/4 in. (19 mm), although attrition measurements can also be performed on larger size materials. Variable and constant rate rotating cylinder abrasimeters designed according to ASTM D4058-96 are readily available. Typically, the material to be tested is placed in drum of the rotating test cylinder and rolled at from 55 to 65 RPM for 35 minutes. Afterwards, the material is removed from test cylinder and screened on a 20-mesh sieve. The percentage (by weight) of the original material sample that remains on the 20-mesh sieve is referred to as the "percent retained." The shaped porous carbon products (e.g., extrudates) and catalysts prepared therefrom typically exhibit a rotating drum attrition index as measured in accordance with ASTM D4058-96 or similar test method such that the percent retained is greater than 85% by weight. A percent retained result of greater than 97% is indicative of materials with exceptional mechanical stability and robust structure particularly desirable for industrial applications.

[0059] Abrasion loss (ABL) is an alternate measurement of the resistance of the shaped porous carbon products (e.g., extrudates) and catalysts prepared therefrom. As with the attrition index, the results of this test method can be used as a measure of fines production during the handling, transportation, and use of the material. Abrasion loss is a measurement of the resistance of a material to attrition wear due to the intense horizontal agitation of the particles within the confines of a 30-mesh sieve. Typically, the material to be tested is first de-dusted on a 20-mesh sieve by gently moving the sieve side-to-side at least 20 times. The de-dusted sample is weighed and then transferred to the inside of a clean, 30-mesh sieve stacked above a clean sieve pan for the collection of fines. The complete sieve stack is then assembled onto a sieve shaker (e.g., RO-Tap RX-29 sieve shaker from W.S. Tyler Industrial Group, Mentor, OH), covered securely and shaken for 30 minutes. The collected fines generated are weighed and divided by the de-dusted sample weight to provide a sample abrasion loss in percent by weight. The shaped porous carbon products (e.g., extrudates) and catalysts prepared therefrom typically exhibit a horizontal agitation sieve abrasion loss of less than 5%. An abrasion loss result of less than 2% is particularly desired for industrial applications.

[0060] The shaped porous carbon products and methods for preparing the shaped porous carbon products include various combinations of features described herein. For example, , the shaped porous carbon product may comprise (a) carbon black and (b) a carbonized binder comprising a carbonization product of a water soluble organic binder, wherein the shaped porous carbon product has a BET specific surface area from 20 m²/g to 500 m²/g, a mean pore diameter greater than 10 nm, a specific pore volume greater than 0.1 cm³/g, a radial piece crush strength greater than 4.4 N/mm (1lb/mm), and a carbon black content of at least 35 wt.%. The shaped porous carbon product may comprise a carbon agglomerate, wherein the shaped porous carbon product has a mean diameter of at least 50 μm, a BET specific surface area from 20 m²/g to 500 m²/g, a mean pore diameter greater than 10 nm, a specific pore volume greater than 0.1 cm³/g, and a radial piece crush strength greater than 4.4 N/mm (1 lb/mm).

[0061] The shaped porous carbon product may also have a low sulfur content. For example, the sulfur

content of the shaped porous carbon product may be no greater than 1 wt.%.

[0062] Features and characteristics including the type of carbon black, the binder, the specific surface area, the specific pore volume, the mean pore diameter, the crush strength, attrition and abrasion resistance and the carbon black content may be independently adjusted or modified within the ranges described herein. Also, shaped porous carbon products may be further defined according to characteristics described herein.

[0063] For instance, the shaped porous carbon product can comprises (a) carbon black and (b) a carbonized binder comprising a carbonization product of a water soluble organic binder and wherein the shaped porous carbon product has a BET specific surface area from 20 m²/g to 500 m²/g, a mean pore diameter greater than 5 nm, a specific pore volume greater than 0.1 cm³/g, a radial piece crush strength greater than 4.4 N/mm (1 lb/mm), a carbon black content of at least 35 wt.%, and a carbonized binder content from 20 wt.% to 50 wt.%.

[0064] Alternatively, a shaped porous carbon product may comprise (a) carbon black and (b) a carbonized binder comprising a carbonization product of a water soluble organic binder, wherein the shaped porous carbon product has a BET specific surface area from 20 m²/g to 500 m²/g, a mean pore diameter greater than 10 nm, a specific pore volume greater than 0.1 cm³/g, and a radial piece crush strength greater than 4.4 N/mm (1 lb/mm), and a carbon black content of at least 35 wt.%, and wherein the shaped porous carbon product has a pore volume measured on the basis of pores having a diameter from 1.7 nm to 100 nm and at least 35% of the pore volume is attributable to pores having a mean pore diameter of from 10 nm to 50 nm.

[0065] Yet another shaped porous carbon product may comprise a carbon agglomerate, wherein the shaped porous carbon product has a mean diameter of at least 50 μm, a BET specific surface area from 20 m²/g to 500 m²/g, a mean pore diameter greater than 10 nm, a specific pore volume greater than 0.1 cm³/g, and a radial piece crush strength greater than 4.4 N/mm (1 lb/mm), and wherein the shaped porous carbon product has a pore volume measured on the basis of pores having a diameter from 1.7 nm to 100 nm and at least 35% of the pore volume is attributable to pores having a mean pore diameter of from 10 nm to 50 nm.

[0066] . The carbon black mixture may optionally be heated during the forming step (e.g., extrusion, pelletizing, pilling, tableting, cold or hot isostatic pressing, calandering, injection molding, 3D printing, drip casting, or other methods) to facilitate the forming of the carbon black mixture into the desired shape.

[0067] The shaped porous carbon black products can also be wash-coated or dip-coated onto other materials to prepare structured composite materials. The shaped porous carbon black products (at least micron-sized) can be domains on heterogeneous, segregated composite materials (e.g., carbon - ZrO₂ composites or carbon domains hosted by large-pore (mm-sized) ceramic foams) as well as layered or structured materials (e.g., carbon black wash-coats onto inert supports such as steatite, plastic or glass balls).

[0068] The shaped porous carbon black product of the invention may be further treated thermally or chemically to alter the physical and chemical characteristics of the shaped porous carbon black product. For example, chemical treatment such as an oxidation may produce a more hydrophilic surface which may provide advantages for preparing a catalyst (improved wetting and dispersion). Oxidation methods are known in the art, see for example U.S. Patents 7,922,805 and 6,471,763. In other embodiments, the

shaped porous carbon black product has been surface treated using known methods for attaching a functional group to a carbon based substrate. See, e.g., WO2002/018929, WO97/47691, WO99/23174, WO99/31175, WO99/51690, WO2000/022051, and WO99/63007. The functional group may be an ionizable group such that when the shaped porous carbon black product is subjected to ionizing conditions, it comprises an anionic or cationic moiety. This embodiment is useful when the shaped porous carbon black product is used as a separation media in chromatography columns and other separation devices.

Catalyst Compositions and Method of Preparation

[0069] The present invention is directed to a method of preparing a catalyst composition comprising the shaped porous carbon product as a catalyst support. The shaped porous carbon products provide effective dispersion and anchoring of catalytically active components or precursors thereof to the surface of the carbon product. The catalyst compositions of the present invention are suitable for use in long term continuous flow operation phase reactions under demanding reaction conditions such as liquid phase reactions in which the shaped porous carbon product is exposed to reactive solvents such as acids and water at elevated temperatures. The catalyst compositions comprising the shaped porous carbon products of the present invention demonstrate operational stability necessary for commodity applications.

[0070] In general, the catalyst compositions prepared by the method of the present invention comprise the shaped porous carbon product as a catalyst support and a catalytically active component or precursor thereof at a surface of the support (external and/or internal surface). In various catalyst compositions of the present invention, the catalytically active component or precursor thereof comprises a noble metal at a surface of the shaped porous carbon product. Typically, the metal(s) at a surface of the catalyst support may constitute from 0.1% to 50% of the total weight of the catalyst.

[0071] In general, the noble metals may be present in various forms (e.g., elemental, metal oxide, metal hydroxides, metal ions, metalates, polyanions, oligomers or colloidal etc.). Typically, however, the noble metals are reduced to elemental form during preparation of the catalyst composition or *in-situ* in the reactor under reaction conditions.

[0072] The noble metal(s) may be deposited on a surface of the shaped porous carbon product according to procedures known in the art including, but not limited to incipient wetness, ion-exchange, deposition-precipitation, coating and vacuum impregnation. When two or more noble metals are deposited on the same support, they may be deposited sequentially or simultaneously. Multiple impregnation steps are also possible (e.g., dual impregnation of the same noble metal under different conditions to increase overall metal loading or tune the metal distribution across the shell). The noble metal(s) deposited on the shaped porous carbon product for the oxidation catalyst may form a shell at least partially covering the surface of the carbon product. In other words, noble metal deposited on the shaped porous carbon product coats external surfaces of the carbon product. The noble metal may penetrate surficial pores of the shaped porous carbon product to form a shell layer ("egg shell") with a thickness of from 10 μm to 400 μm . The shell may be produced sub-surface to produce a 10 μm to 400 μm sub-surface band containing the catalytically active noble metals ("egg yolk"). Also structured shells featuring different noble metal distributions across the shell for the various metals are possible.

[0073] Following noble metal deposition, the catalyst composition prepared by the method of the present invention is optionally dried, for example, at a temperature of at least 50°C for a period of time of at least 1 hour. Alternatively, the drying may be conducted in a continuous or staged manner where independently controlled temperature zones (e.g., 60°C, 80°C, and 120°C) are utilized. Typically, drying is initiated below

the boiling point of the solvent, e.g., 60°C and then increased. The catalyst is dried under sub-atmospheric or atmospheric pressure conditions. The catalyst may be reduced after drying (e.g., by flowing 5% H₂ in N₂ at 350 °C for 3 hours). Still further, the catalyst may be calcined, for example, at a temperature of at least 200°C for a period of time (e.g., at least 3 hours).

[0074] The catalyst composition prepared by the method of the present invention is prepared by depositing the catalytically active component or precursor thereof subsequent to forming the shaped porous carbon product (i.e., depositing directly on a surface of the shaped porous carbon product). The catalyst composition of the present invention may be prepared by contacting the shaped porous carbon product with a solubilized noble metal complex or combination of solubilized noble metal complexes. The heterogeneous mixture of solid and liquids can then be stirred, mixed and/or shaken to enhance the uniformity of dispersion of the catalyst, which, in turn, enables the more uniform deposition of metal(s) on the surface of the support upon removal of the liquids. Following deposition, the noble metal complex(es) on the shaped porous carbon products are heated and reduced under a reducing agent such as a hydrogen containing gas (e.g., forming gas 5% H₂ and 95% N₂). The temperature at which the heating is conducted generally ranges from 150°C to 600°C. Heating is typically conducted for a period of time ranging from 1 hour to 5 hours. Reduction may also be carried in the liquid phase. For example, catalyst compositions can be treated in a fixed bed with the liquid containing a reducing agent pumped through the static catalyst.

[0075] The catalyst compositions prepared by the method of the present invention comprising the shaped porous carbon product as a catalyst support can be deployed in various reactor formats, particularly those suited liquid phase medium such as batch slurry, continuous slurry-based stirred tank reactors, cascade of stirred tank reactors, bubble slurry reactor, fixed beds, ebulated beds and other known industrial reactor formats.

Catalytic Processes

[0076] The catalyst compositions prepared by the method of the present invention are useful for various catalytic conversions including oxidations, reductions, dehydrations, hydrogenations and other known transformations using appropriate active metals formulations and which can be conducted in gaseous or liquid medium.

[0077] Processes comprise contacting a liquid medium comprising the reactant with a catalyst composition prepared by them method of the present invention. Typically, the catalyst composition has superior mechanical strength (e.g., mechanical piece crush strength and/or radial piece crush strength) and is stable to the continuous flow of the liquid medium and reaction conditions for at least 500 hours without substantial loss in catalytic productivity, selectivity, and/or yield.

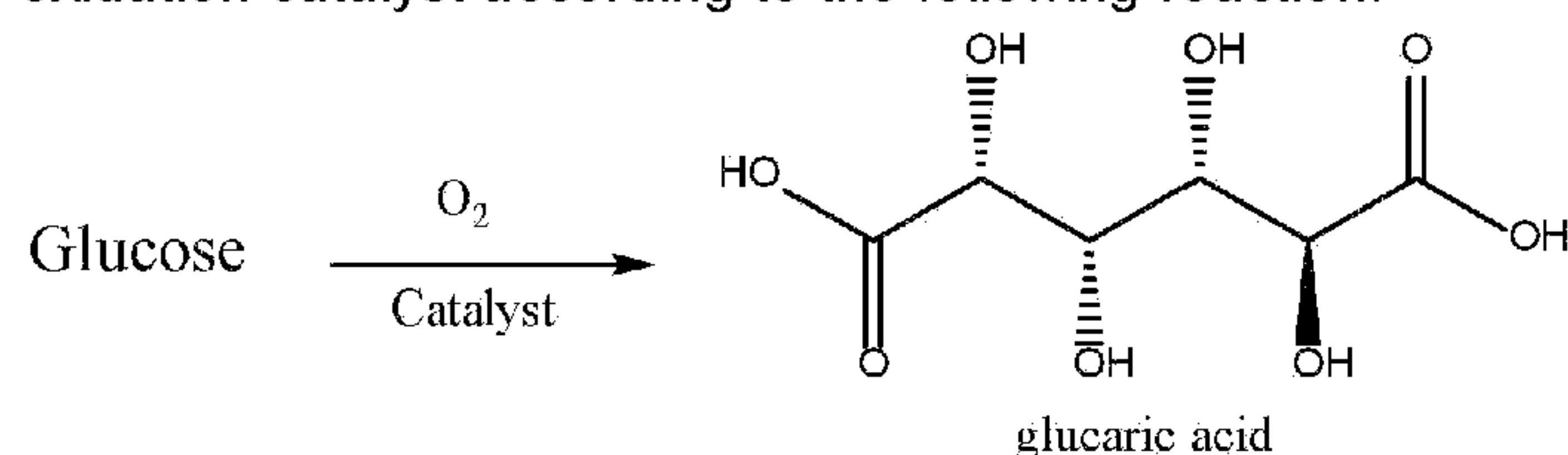
[0078] In addition, it has been surprisingly discovered that the catalyst compositions prepared by the method of the present invention are highly productive and selective catalysts for a certain of chemical transformations such as the conversion of highly functionalized and/or non-volatile molecules including, but not limited to biorenewably-derived molecules and intermediates for commodity applications.

Catalytic Oxidation

[0079] One series of chemical transformations that the catalyst compositions prepared by the method of

the present invention are suited for is the selective oxidation of a hydroxyl group to a carboxyl group in a liquid or gaseous reaction medium. For example, one series of chemical transformations that the catalyst compositions of the present invention are especially suited for is the selective oxidation an aldose to an aldarcic acid. Accordingly, catalyst compositions of the present invention as described herein can be utilized as oxidation catalysts. Aldoses include, for example, pentoses and hexoses (i.e., C-5 and C-6 monosaccharides). Pentoses include ribose, arabinose, xylose, and lyxose, and hexoses include glucose, allose, altrose, mannose, gulose, idose, galactose, and talose. Accordingly, in various embodiments, the present invention is also directed to a process for the selective oxidation of an aldose to an aldarcic acid comprising reacting the aldose with oxygen in the presence of a catalyst composition as described herein to form the aldarcic acid. Typically, the catalyst composition comprises at least platinum as a catalytically active component.

[0080] The catalyst compositions prepared by the method of the present invention have been found to be especially selective for the oxidation of the glucose to glucaric acid. Accordingly, a process for the selective oxidation of glucose to glucaric acid comprising reacting the aldose with oxygen in the presence of a catalyst composition as described herein to form glucaric acid is described. U.S. Patent No. 8,669,397 discloses various catalytic processes for the oxidation of glucose to glucaric acid. In general, glucose may be converted to glucaric acid in high yield by reacting glucose with oxygen (e.g., air, oxygen-enriched air, oxygen alone, or oxygen with other constituents substantially inert to the reaction) in the presence of an oxidation catalyst according to the following reaction:



The oxidation can be conducted in the absence of added base (e.g., KOH) or where the initial pH of the reaction medium and/or the pH of reaction medium at any point in the reaction is no greater than 7. The initial pH of the reaction mixture is the pH of the reaction mixture prior to contact with oxygen in the presence of an oxidation catalyst. In fact, catalytic selectivity can be maintained to attain glucaric acid yield in excess of 30%. The absence of added base advantageously facilitates separation and isolation of the glucaric acid, thereby providing a process that is more amenable to industrial application, and improves overall process economics by eliminating a reaction constituent. The "absence of added base" as used herein means that base, if present (for example, as a constituent of a feedstock), is present in a concentration which has essentially no effect on the efficacy of the reaction; i.e., the oxidation reaction is being conducted essentially free of added base. The oxidation reaction can also be conducted in the presence of a weak carboxylic acid, such as acetic acid, in which glucose is soluble. The term "weak carboxylic acid" as used herein means any unsubstituted or substituted carboxylic acid having a pKa of at least 3.5, more particularly, is selected from among unsubstituted acids such as acetic acid, propionic acid or butyric acid, or mixtures thereof.

[0081] The oxidation reaction may be conducted under increased oxygen partial pressures and/or higher oxidation reaction mixture temperatures, which tends to increase the yield of glucaric acid when the reaction is conducted in the absence of added base or at a pH below 7. Typically, the partial pressure of oxygen is at least 15 pounds per square inch absolute (psia) (104 kPa). The partial pressure of oxygen may be up to 1,000 psia (6895 kPa). Generally, the temperature of the oxidation reaction mixture is at least 40°C, such as from 40°C to 200°C. Surprisingly, the catalyst compositions prepared by the method of the present invention permit glucose oxidation at elevated temperatures (e.g., from 100°C to 160°C) without heat degradation of the catalyst. In particular, reactor formats such as a fixed bed reactor which can provide a relatively high liquid throughput in combination with the catalyst compositions comprising the shaped porous carbon product comprising carbon black have been found to permit oxidation at

temperatures in excess of 140°C (e.g., 140° to 150°C).

[0082] Oxidation of glucose to glucaric acid can also be conducted in the absence of nitrogen as an active reaction constituent. Some processes employ nitrogen compounds such as nitric acid as an oxidant. The use of nitrogen in a form in which it is an active reaction constituent, such as nitrate or nitric acid, results in the need for NO_x abatement technology and acid regeneration technology, both of which add significant cost to the production of glucaric acid from these known processes, as well as providing a corrosive environment which may deleteriously affect the equipment used to carry out the process. By contrast, for example, in the event air or oxygen-enriched air is used in the oxidation reaction of the present invention as the source of oxygen, the nitrogen is essentially an inactive or inert constituent. Thus, an oxidation reaction employing air or oxygen-enriched air is a reaction conducted essentially free of nitrogen in a form in which it would be an active reaction constituent.

[0083] Applicants have discovered that oxidation catalyst compositions prepared by the method of the present invention provide unexpectedly greater selectivity and yield for producing glucaric acid from glucose when compared to similar catalysts comprising similar support materials such as activated carbon. In particular, applicants have unexpectedly found that enhanced selectivity and yield for glucaric acid can be achieved by use of an oxidation catalyst comprising platinum and gold at a surface of the shaped porous carbon product (i.e., at a surface of the catalyst support).

[0084] The enhanced glucaric acid yield is typically at least 30% (e.g., from 35% to 65%). Further, the enhanced glucaric acid selectivity is typically at least 70%.

[0085] The catalytically active components or precursors thereof comprising platinum and gold may be in the form described in U.S. Patent Application Publication 2011/0306790. This publication describes various oxidation catalysts comprising a catalytically active component comprising platinum and gold, which are useful for the selective oxidation of compositions comprised of a primary alcohol group and at least one secondary alcohol group (e.g., glucose).

[0086] An oxidation catalyst composition prepared by a method of the present invention may comprise particles of gold in the form of a gold-containing alloy and particles consisting essentially of platinum (0) as the catalytically active components on a surface of the catalyst support. Typically, the total noble metal loading of the catalyst composition is 10 wt.% or less.

[0087] In order to oxidize glucose to glucaric acid, a sufficient amount of the catalytically active component must be present relative to the amount of reactant (i.e., glucose). Accordingly, for the oxidation of glucose to glucaric acid as described herein where the catalytically active component comprises platinum, typically the mass ratio of glucose to platinum is from 10:1 to 1000:1.

[0088] In the method of the present invention, the gold component of the catalyst may be typically added to the shaped porous carbon product as a solubilized constituent to enable the formation of a uniform suspension. A base is then added to the suspension in order to create an insoluble gold complex which can be more uniformly deposited onto the support. For example, the solubilized gold constituent is provided to the slurry as gold salt, such as HAuCl₄. Upon creation of a well dispersed, heterogeneous mixture, a base is added to the slurry to form an insoluble gold complex which then deposits on the surface of the shaped porous carbon product. Although any base which can affect the formation of an insoluble gold complex is useable, bases such as KOH, NaOH are typically employed. It may be desirable, though not required, to collect the shaped porous carbon product on which has been deposited the insoluble gold complex prior to adding the platinum-containing constituent, which collection can readily be accomplished by any of a variety of means known in the art such as, for example, centrifugation. The collected solids may optionally be

washed and then may be heated to dry. Heating may also be employed so as to reduce the gold complex on the support to gold (0). Heating may be conducted at temperatures ranging from 60°C (to dry) up to 500°C (at which temperature the gold can be effectively reduced). The heating step may be conducted in the presence of a reducing or oxidizing atmosphere in order to promote the reduction of the complex to deposit the gold onto the support as gold (0). Heating times vary depending upon, for example, the objective of the heating step and the decomposition rate of the base added to form the insoluble complex, and the heating times can range from a few minutes to a few days. More typically, the heating time for the purpose of drying ranges from 2 to 24 hours and for reducing the gold complex is on the order of 1 to 4 hours.

[0089] The concentration of the shaped porous carbon product in the slurry can be in the range of 1 to 100 g of solid/liter of slurry.

[0090] Mixing of the slurry containing the soluble gold-containing compound is continued for a time sufficient to form at least a reasonably uniform suspension. Appropriate times can range from minutes to a few hours. After addition of the base to convert the gold-containing compound to an insoluble gold-containing complex, the uniformity of the slurry should be maintained for a time sufficient to enable the insoluble complex to be formed and deposit on the shaped porous carbon product. The time can range from a few minutes to several hours.

[0091] Platinum can be added to the shaped porous carbon product or slurry thereof after deposition of gold onto the shaped porous carbon product or after heat treatment to reduce the gold complex on the support to gold (0). Alternatively, the platinum may be added to the shaped porous carbon product or slurry thereof prior to the addition of the solubilized gold compound provided the platinum present on the support is in a form that will not be re-dissolved upon the addition of base used to promote the deposition of gold onto the support. The platinum is typically added as a solution of a soluble precursor or as a colloid. Platinum may be added as a compound selected from the group of platinum (II) nitrate, platinum(IV) nitrate, platinum oxynitrate, platinum (II) acetylacetonate (acac), tetraamineplatinum (II) nitrate, tetraamineplatinum (II) hydrogenphosphate, tetraamineplatinum (II) hydrogencarbonate, tetraamineplatinum (II) hydroxide, H_2PtCl_6 , PtCl_4 , Na_2PtCl_4 , K_2PtCl_4 , $(\text{NH}_4)_2\text{PtCl}_4$, $\text{Pt}(\text{NH}_3)_4\text{Cl}_2$, mixed $\text{Pt}(\text{NH}_3)_x\text{Cl}_y$, $\text{K}_2\text{Pt}(\text{OH})_6$, $\text{Na}_2\text{Pt}(\text{OH})_6$, $(\text{NMe}_4)_2\text{Pt}(\text{OH})_6$, and $(\text{EA})_2\text{Pt}(\text{OH})_6$ where EA = ethanolamine. More preferred compounds include platinum(II) nitrate, platinum(IV) nitrate, platinum(II) acetylacetonate (acac), tetraamine platinum(II) hydroxide, K_2PtCl_4 , and $\text{K}_2\text{Pt}(\text{OH})_6$.

[0092] Subsequent to the addition of the platinum compound, the support slurry and platinum-containing compound is dried. Drying may be conducted at room temperature or at a temperature up to 120°C. The drying step may be conducted for a period of time ranging from a few minutes to a few hours. Typically, the drying time is in the range of 6 hours to 24 hours. The drying can also be done with continuous or staged temperature increase from 60°C to 120°C on a band calciner or belt dryer (which is preferred for commercial applications).

[0093] After drying the support having the platinum compound deposited thereon, it is subjected to at least one thermal treatment in order to reduce platinum deposited as platinum (II) or platinum (IV) to platinum (0). The thermal treatment(s) can be conducted in air or in any reducing or oxidizing atmosphere. The thermal treatment(s) may be conducted under a forming gas atmosphere. Alternatively, a liquid reducing agent may be employed to reduce the platinum; for example, hydrazine or formaldehyde or formic acid or salts thereof (e.g., sodium formate) or NaH_2PO_2 may be employed to effect the requisite reduction of the platinum. The atmosphere under which the thermal treatment is conducted is dependent upon the platinum compound employed, with the objective being substantially converting the platinum on the support to platinum (0).

[0094] The temperatures at which the thermal treatment(s) is (are) conducted generally range from 150°C to 600°C. The thermal treatment is typically conducted for a period of time ranging from 1 hour to 8 hours.

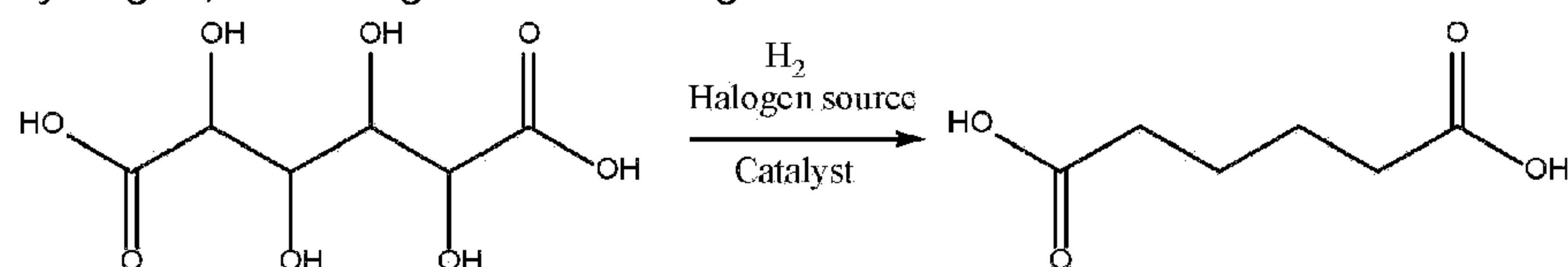
[0095] The metal(s) deposited on the shaped porous carbon product for the oxidation catalyst may form a shell at least partially covering the surface of the carbon product. In other words, metal deposited on the shaped porous carbon product coats external surfaces of the carbon product. The metal may penetrate surficial pores of the shaped porous carbon product to form a shell layer ("egg shell") with a thickness of from 10 μm to 400 μm . The shell may be produced sub-surface to produce a 10 μm to 400 μm sub-surface band containing the catalytically active metals ("egg yolk").

Catalytic Hydrodeoxygenation

[0096] One series of chemical transformations that the catalyst compositions prepared by the method of the present invention are suited for is the hydrodeoxygenation of carbon-hydroxyl groups to carbon-hydrogen groups in a liquid or gaseous reaction medium. For example, one series of chemical transformation that the catalyst compositions of the present invention are especially suited for is the selective halide-promoted hydrodeoxygenation of an aldaric acid or salt, ester, or lactone thereof to a dicarboxylic acid. Accordingly, catalyst compositions prepared by the method of the present invention as described herein can be utilized as hydrodeoxygenation catalysts. As such, also described here is a process for the selective halide promoted hydrodeoxygenation of an aldaric acid comprising reacting the aldaric acid or salt, ester, or lactone thereof with hydrogen in the presence of a halogen-containing compound and a catalyst composition prepared by a method of the present invention to form a dicarboxylic acid.

[0097] The catalyst compositions prepared by the method of the present invention have been found to be especially selective for halide-promoted hydrodeoxygenation of glucaric acid or salt, ester, or lactone thereof to adipic acid. U.S. Patent No. 8,669,397, referenced above, describes the chemocatalytic processes for the hydrodeoxygenation of glucaric acid to adipic acid.

[0098] Adipic acid or salts and esters thereof may be prepared by reacting, in the presence of a hydrodeoxygenation catalyst and a halogen source, glucaric acid or salt, ester, or lactone thereof, and hydrogen, according to the following reaction:



[0099] In the above reaction, glucaric acid or salt, ester, or lactone thereof is converted to an adipic acid product by catalytic hydrodeoxygenation in which carbon-hydroxyl groups are converted to carbon-hydrogen groups. The catalytic hydrodeoxygenation may be hydroxyl-selective wherein the reaction is completed without substantial conversion of the one or more other non-hydroxyl functional group of the substrate.

[0100] The halogen source may be in a form selected from the group consisting of ionic, molecular, and mixtures thereof. Halogen sources include hydrohalic acids (e.g., HCl, HBr, HI and mixtures thereof; preferably HBr and/or HI), halide salts, (substituted or unsubstituted) alkyl halides, or molecular (diatomic)

halogens (e.g., chlorine, bromine, iodine or mixtures thereof; preferably bromine and/or iodine). The halogen source may be in diatomic form, hydrohalic acid, or halide salt and, more preferably, diatomic form or hydrohalic acid. The halogen source may be a hydrohalic acid, in particular hydrogen bromide.

[0101] Generally, the molar ratio of halogen to the glucaric acid or salt, ester, or lactone thereof is equal to or less than 1. The mole ratio of halogen to the glucaric acid or salt, ester, or lactone thereof may typically be from 1:1 to 0.1:1.

[0102] Generally, the reaction allows for recovery of the halogen source and catalytic quantities (where molar ratio of halogen to the glucaric acid or salt, ester, or lactone thereof is less than 1) of halogen can be used, recovered and recycled for continued use as a halogen source.

[0103] Generally, the temperature of the hydrodeoxygenation reaction mixture is at least 20°C, such as from 20°C to 250°C. Typically, the partial pressure of hydrogen is at least 25 psia (172 kPa), such as from 25 psia (172 kPa) to 2500 psia (17237 kPa).

[0104] The hydrodeoxygenation reaction may be conducted in the presence of a solvent. Solvents suitable for the selective hydrodeoxygenation reaction include water and carboxylic acids, amides, esters, lactones, sulfoxides, sulfones and mixtures thereof.

[0105] Applicants have discovered that hydrodeoxygenation catalyst compositions prepared by the method of the present invention provide enhanced selectivity and yield for producing adipic acid. In particular, applicants have unexpectedly found that enhanced selectivity and yield for adipic acid can be achieved by use of a catalyst composition prepared by the method.

[0106] The catalytically active component or precursor thereof may include noble metals selected from the group consisting of ruthenium, rhodium, palladium, platinum, and combinations thereof. The hydrodeoxygenation catalyst may comprise two or more noble metals. The first metal may be selected from the group of platinum, rhodium and palladium, and the second metal may be selected from the group consisting of ruthenium, rhodium, palladium, platinum, and gold. For example, the first metal may be platinum and the second metal may be rhodium. The platinum to rhodium molar ratio of the catalyst composition prepared in accordance with the present invention may be in the range of from 3:1 to 1:2.

[0107] The noble metal(s) deposited on the shaped porous carbon product for the hydrodeoxygenation catalyst may form a shell at least partially covering the surface of the carbon product. In other words, metal deposited on the shaped porous carbon product coats external surfaces of the carbon product. The metal may penetrate surficial pores of the shaped porous carbon product to form a shell layer ("egg shell") with a thickness of from 10 µm to 400 µm. The shell may be produced sub-surface to produce a 10 µm to 400 µm sub-surface band containing the catalytically active metals ("egg yolk").

Hydrodeoxygenation of 1,2,6-Hexanetriol

[0108] Another chemical transformation that the catalyst compositions prepared by the method of the present invention is advantageous for is the selective hydrodeoxygenation of 1,2,6-hexanetriol to 1,6-hexanediol (HDO) and 1,2,5,6-hexanetetraol to 1,6-HDO). The catalytically active component of the catalyst composition may comprise platinum. The catalytically active component of the catalyst composition may comprise platinum and at least one metal (M2) selected from the group of molybdenum, lanthanum, samarium, yttrium, tungsten, and rhenium. The catalytically active component of the catalyst composition

may comprise platinum and tungsten.

[0109] Typically, the total weight of metal(s) is from 0.1% to 10% of the total weight of the catalyst. The total weight of metal of the catalyst may be less than 4%. The molar ratio of platinum to (M2) may vary, for example, from 20:1 to 1:10.

[0110] Typically, the conversion of 1,2,6-hexanetriol to HDO is conducted at a temperature in the range of 60° C to 200° C and a partial pressure of hydrogen in the range of 200 psig (1480 kPa) to 2000 psig (13891 kPa).

Catalytic Amination of 1,6-Hexanediol

[0111] Furthermore, the catalyst compositions prepared by the method of the present invention are also useful for the selective amination of 1,6-hexanediol (HDO) to 1,6-hexamethylenediamine (HMDA). The catalytically active component of the catalyst composition may comprise ruthenium.

[0112] The catalytically active component of the catalyst composition may comprise ruthenium and optionally a second metal such as rhenium or nickel. One or more other d-block metals, one or more rare earth metals (e.g., lanthanides), and/or one or more main group metals (e.g., Al) may also be present in combination with ruthenium and with ruthenium and rhenium combinations. The catalytically active phase may consist essentially of ruthenium and rhenium. Typically, the total weight of metal(s) is from 0.1% to 10% of the total weight of the catalyst composition.

[0113] When the catalysts prepared by the method of the present invention comprise ruthenium and rhenium in combination, the molar ratio of ruthenium to rhenium is important. A by-product of processes for converting HDO to HMDA is pentylamine. Pentylamine is an off path by-product of the conversion of HDO to HMDA that cannot be converted to HMDA or to an intermediate which can, on further reaction in the presence of the catalysts of the present invention, be converted to HMDA. However, the presence of too much rhenium can have an adverse effect on the yield of HMDA per unit area time (commonly known as space time yield, or STY). Therefore, the molar ratio of ruthenium:rhenium should be maintained in the range of from 20:1 to 4:1. A ruthenium:rhenium molar ratio from 8:1 to 4:1 may produce HMDA in at least 25% yield with an HMDA/pentylamine ratio of at least 20:1.

[0114] HDO is converted to HMDA by reacting HDO with an amine, e.g., ammonia, in the presence of the catalysts of the present invention. Generally, the amine may be added to the reaction in the form of a gas or liquid. Typically, the molar ratio of ammonia to HDO is at least 40:1, or it may be in the range of from 40:1 to 5:1. The reaction of HDO with amine in the presence of catalyst composition prepared by a method of the present invention is carried out at a temperature less than or equal to 200°C.

[0115] Generally, the reaction is conducted at a pressure not exceeding 1500 psig (10443 kPa). The reaction pressure may be in the range of 200 psig (1480 kPa) to 1500 psig (10443 kPa). The disclosed pressure ranges may include the pressure of NH₃ gas and an inert gas, such as N₂. The pressure of NH₃ gas may be in the range of 50-150 psig (446-1135 kPa) and an inert gas, such as N₂ is in the range of 700 psig (4927 kPa) to 1450 psig (10098 kPa).

[0116] The catalyst may be contacted with HDO and ammonia at a temperature in the range of 100°C to 180°C and a pressure in the range of 200 psig (1480 kPa) to 1500 psig (10443 kPa). The disclosed pressure ranges may include the pressure of NH₃ gas and an inert gas, such as N₂. The pressure of NH₃

gas may be in the range of 50-150 psig (446-1135 kPa) and an inert gas, such as N₂ is in the range of 500 psig (3548 kPa) to 1450 psig (10098 kPa).

[0117] The process may be carried out in the presence of hydrogen. Typically, where the HDO and amine are reacted in the presence of hydrogen and the catalyst of the present invention, the hydrogen partial pressure is equal to or less than 100 psig (790 kPa).

[0118] The conversion of HDO to HMDA can also be conducted in the presence of a solvent. Solvents suitable for use in conjunction with the conversion of HDO to HMDA in the presence of the catalysts prepared by the method of the present invention may include, for example, water, alcohols, esters, ethers, ketones, or mixtures thereof.

[0119] The chemocatalytic conversion of HDO to HMDA is likely to produce one or more by-products such as, for example, pentylamine and hexylamine. By-products which are subsequently convertible to HMDA by further reaction in the presence of catalysts of the present invention are considered on-path by-products. Other by-products such as, for example, pentylamine and hexylamine are considered off path by-products for the reasons above discussed. At least 20% of the product mixture resulting from a single pass reaction of HDO with amine (e.g., ammonia) in the presence of the catalysts of the present invention is HMDA.

[0120] The product mixture may be separated into one or more products by any suitable methods known in the art. The product mixture can be separated by fractional distillation under subatmospheric pressures. For example, HMDA can be separated from the product mixture at a temperature between 180°C and 220°C. The HDO may be recovered from any remaining other products of the reaction mixture by one or more conventional methods known in the art including, for example, solvent extraction, crystallization or evaporative processes. The on-path by-products can be recycled to the reactor employed to produce the product mixture or, for example, supplied to a second reactor in which the on path by-products are further reacted with ammonia in the presence of a catalysts prepared by a method of the present invention to produce additional HMDA.

[0121] The following non-limiting examples are provided to further illustrate the present invention.

EXAMPLES

[0122] Surface areas were determined from nitrogen adsorption data using the BET method as described in S. Brunauer, P.H. Emmett, E. Teller, J. Am. Chem. Soc. 1938, 60, 309-331, and ASTM D3663-03(2008) Standard Test Method for Surface Area of Catalysts and Catalyst Carriers. Mean pore diameters and pore volumes were determined in accordance with the procedures described in E.P. Barrett, L.G. Joyner, P. P. Halenda, J. Am. Chem. Soc. 1951, 73, 373-380, and ASTM D4222-03(2008) Standard Test Method for Determination of Nitrogen Adsorption and Desorption Isotherms of Catalysts and Catalyst Carriers by Static Volumetric Measurements.

[0123] Mercury Porosimetry measurements were conducted using an AutoPore V Mercury Porosimeter from Micromeritics Instrument Corporation. A suitable amount of carbon extrudates were loaded into an appropriate penetrometer and mercury intrusion was measured in two sequential stages: low pressure analysis (0 to 50 psia) (0 to 345 kPa) followed by high pressure analysis (50 to 33,000 psia) (345 to 226527 kPa). A total of 712 data points were collected across the whole range of pressure with a contact angle of 154.0°.

[0124] Radial piece crush strength measurements were conducted according to ASTM D6175 - 03(2013) Standard Test Method for Radial Crush Strength of Extruded Catalyst and Catalyst Carrier Particles using a press apparatus equipped with a Dillon GS100 Digital Force Gauge. Mean radial piece crush strength is the average value of independent measurements of at least 10 different extrudate pellets.

[0125] Single piece crush strength were conducted according to ASTM D4179 - 03(2013) Standard Test Method for Single Pellet Crush Strength of Formed Catalysts and Catalyst Carriers using a press apparatus equipped with a Dillon GS100 Digital Force Gauge. Single piece crush strength is the average value of independent measurements of at least 10 different extrudate pellets.

Example 1. Preparation of Carbon Black Extrudates

[0126] 36.4 g of carbon black powder (Cabot Vulcan XC72, 224 m²/g) was added in multiple portions to a heated (overnight at 80°C) aqueous solution (136.5 g) containing 24.3 wt.% Cerelese Dextrose from Ingredion and 4.7 wt.% hydroxyethylcellulose from Sigma-Aldrich (SKU 54290, viscosity 80-125 cP (80-125 mPa. s), 2% in H₂O (20°C)). The mixture was stirred well using a spatula to produce a paste. This paste was loaded into a syringe and the material was extruded into spaghetti-like strings with 1.5 mm diameter. After drying in a 70°C oven for 5 hours under a dry air purge, these strings were cut into small pieces 1.0 cm long. Then they were treated at 350°C for 2 hours with 10°C/min temperature ramp rate under continuous N₂ flow to carbonize the binder and produce a carbon black extrudate.

[0127] The properties of the resultant extrudate are show in Table 1.

Table 1.

	BET Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)
Extrudate of Cabot Vulcan XC72	110	110	0.15

Example 2. Characterization of Component Carbon Black Powder

[0128] The properties of various carbon black powders utilized in the shaped porous carbon black products were characterized.

A. Physical Properties of Carbon Black Powders

[0129] The BET specific surface area, specific mean pore diameter and specific pore volume of these carbon black powder starting materials were determined using the methods described above, and are provided in Table 2.

Table 2.

Carbon Black	BET Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)
Asbury 5365R	34	143	0.11
Asbury 5353R	35	186	0.14

Carbon Black	BET Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)
Asbury 5345R	35	207	0.11
Timcal Ensaco 250G	64	140	0.24
Asbury 5348R	65	220	0.31
Asbury 5358R	67	213	0.34
Asbury 5346R	80	145	0.28
Cabot Monarch 570	102	138	0.30
Orion HP 160	158	208	0.87
Cabot Monarch 700	181	121	0.38
Cabot Vulcan XC72	224	161	0.43

B. Catalytic Performance

[0130] The carbon black powders were evaluated as catalyst support material in an oxidation reaction for converting glucose to glucaric acid as described below.

(i) Oxidation of Glucose to Glucaric Acid (Protocol 1)

[0131] Suitably concentrated aqueous solutions of Me₄NAuO₂ and PtO(NO₃) were added together to carbon black powders by incipient wetness impregnation and agitated to impregnate the supports. The samples were dried in an oven at 70°C overnight, and reduced at 350°C under a forming gas (5% H₂ and 95% N₂) atmosphere for 4 hours with 2 °C/min temperature ramp rate to produce catalysts with a composition of 2.0wt% Au and 2.0wt% Pt. By using other carbon black supports, Au and Pt precursors, and adjusting amount of Au and Pt in solution, different catalysts with various Au and Pt loadings on a variety of commercial carbon black powders or particles from extrudates were prepared in a similar manner.

[0132] These catalysts were tested for glucose oxidation using the following testing protocol. Catalyst (8 mg) was weighed into a glass vial insert followed by addition of an aqueous glucose solution (250 µl of 0.55 M). The glass vial insert was loaded into a reactor and the reactor was closed. The atmosphere in the reactor was replaced with oxygen and pressurized to 75 psig (618 kPa) at room temperature. Reactor was heated to 110°C and maintained at the respective temperature for 2 hours while vials were shaken. After that, shaking was stopped and reactor was cooled to 40°C. Pressure in the reactor was then slowly released. The glass vial insert was removed from the reactor and centrifuged. The solution was diluted with deionized water and analyzed by ion chromatography to determine the yield of glucaric acid. Selectivity is defined as 100% x (glucaric acid) / (sum of glucaric acid and all off-pathway species). Off-pathway species that cannot be converted to glucaric acid include 2-ketogluconic acid, 3-ketogluconic acid, 4-ketogluconic acid, 5-ketogluconic acid, trihydroxyglutaric acid, tartaric acid, tartronic acid and oxalic acid. On-pathway species include glucose, gluconic acid, guluronic acid and glucuronic acid. On-pathway species are not used in the selectivity calculation because these intermediates can be partially converted to glucaric acid and are not considered off-pathway. Results are presented in Table 3.

Table 3.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Glucaric Acid Yield (%)	Selectivity (%)
Asbury 5302	211	99	0.29	33	76
Asbury 5303	219	79	0.23	38	77
Asbury 5368	303	102	0.54	34	78
Asbury 5379	271	163	0.83	33	75

(ii) Oxidation of Glucose to Glucaric Acid (Protocol 2)

[0133] Suitably concentrated aqueous solutions of K₂Pt(OH)₆ and CsAuO₂ were added together to carbon black powders by incipient wetness impregnation and agitated to impregnate the supports. The samples were dried in an oven at 40°C overnight, and reduced at 250°C under a forming gas (5% H₂ and 95% N₂) atmosphere for 3 hours with 5 °C/min temperature ramp rate. The final catalysts were washed with deionized water and finally dried at 40°C to produce catalysts with a composition of 2.44 wt.% Pt and 2.38 wt.% Au. By using other carbon black supports, Au and Pt precursors, and adjusting amount of Au and Pt in solution, different catalysts with various Au and Pt loadings on a variety of commercial carbon black powders or particles from extrudates were prepared in a similar manner.

[0134] These catalysts were tested for glucose oxidation using the following testing protocol. Catalyst (10 mg) was weighed into a glass vial insert followed by addition of an aqueous glucose solution (250 µl of 0.55 M). The glass vial insert was loaded into a reactor and the reactor was closed. The atmosphere in the reactor was replaced with oxygen and pressurized to 75 psig (618 kPa) at room temperature. Reactor was heated to 90°C and maintained at the respective temperature for 5 hours while vials were shaken. After that, shaking was stopped and reactor was cooled to 40°C. Pressure in the reactor was then slowly released. The glass vial insert was removed from the reactor and centrifuged. The solution was diluted with deionized water and analyzed by ion chromatography to determine the yield of glucaric acid and the selectivity as defined herein. Results are presented in Table 4.

Table 4.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Glucaric Acid Yield (%)	Selectivity (%)
Asbury 5365R	34	143	0.11	42	78
Asbury 5353R	35	186	0.14	38	82
Asbury 5345R	35	207	0.11	44	76
Timcal Ensaco 250G	64	140	0.24	60	83
Asbury 5348R	65	220	0.31	54	78
Asbury 5358R	67	213	0.34	52	77

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Glucaric Acid Yield (%)	Selectivity (%)
Asbury 5346R	80	145	0.28	48	75
Cabot Monarch 570	102	138	0.3	37	75
Orion HP 160	158	208	0.87	35	77
Cabot Monarch 700	181	121	0.38	42	77
Cabot Vulcan XC72	224	161	0.43	46	77

Example 3. Preparation of Shaped Porous Carbon Black Product Using a Variety of Carbon Black Powders and Binders

[0135] By using other carbon black powders and carbohydrate binders, different carbon black extrudates were prepared as described in Example 1. Other carbon black powders included, but were not limited to, Orion carbon HI-BLACK 40B2, Orion HI-BLACK 50LB, Orion Hi-Black 50L, Orion HP-160, Orion Carbon HI-BLACK N330, Timcal Ensaco 150 G, Timcal Ensaco 250 G, Timcal Ensaco 260G, Timcal Ensaco 250P, Cabot Vulcan XC72R, Cabot Monarch 120, Cabot Monarch 280, Cabot Monarch 570, Cabot Monarch 700, Asbury 5365R, Asbury 5353R, Asbury 5345R, Asbury 5352, Asbury 5374, Asbury 5348R, Asbury 5358R, Sid Richardson SC159, Sid Richardson SR155. Other carbohydrate binders included, but were not limited to, Cargill Clearbrew 60/44 IX (80% Carbohydrate), Casco Lab Fructose 90 (70% Carbohydrate) and Molasses (80% Carbohydrate). Formulations with these variations yielded illustrative examples of the shaped carbon product of the present invention. The properties of some of these embodiments are described in more detail below.

Example 4. Crush Testing of Carbon Black Extrudates

[0136] Extrudate pellets (Nos. 1-8 below), having an approximately 1.5 mm diameter, were prepared according to the method described in Example 1 except that the final pyrolysis times and temperatures were varied as listed in Table 5 in the extrudate description column. After the pyrolysis step the extrudates were cut to the sizes ranging from 2-6 mm in length. The percentage of carbonized binder (after pyrolysis) present in the shaped carbon product was determined by mass balance [i.e., $[(\text{Weight}_{\text{Shaped Carbon Product}} - \text{Weight}_{\text{CarbonBlack (in formulation)}}) / \text{Weight}_{\text{Shaped Carbon Product}}] \times 100$]. Total binder content after pyrolysis (i.e., total carbonized binder) varied from 15-50 wt.%.

[0137] Additional extrudate pellets (Nos. 9-11 below) were prepared accordingly to the following procedure. Approximately 24.0 g of carbon black powder (Timcal Ensaco 250G, 65 m²/g) was added in multiple portions to an aqueous solution (100.0 g) containing 25.0 wt.% Cerelease Dextrose from Ingredion. The mixture was stirred well using a spatula to produce a paste. This paste was loaded into a syringe and the material was extruded into spaghetti-like strings with a 1.5 mm diameter. After drying in a 100°C oven for 3 hours under a dry air purge, these strings were cut into smaller pieces (2-6mm lengths). Then they were treated at one of the following conditions to produce carbon black extrudates: (1) 250°C for 3 hours with

10°C/min temperature ramp rate under continuous N₂ flow; (2) 800°C for 4 hours with 10°C/min temperature ramp rate under continuous N₂ flow; (3) 200°C for 3 hours with 10°C/min temperature ramp rate in air. Binder content varied from 15 to 50 wt.%. Table 5 presents the crush strength data for the extrudates prepared.

Table 5.

No.	Extrudate Description	Mean Mechanical Piece Crush Strength (lb)	Mean Radial Piece Crush Strength (N/mm)	Mean Radial Piece Crush Strength (lb/mm)	Carbonized Binder Content (wt.%)
1	Cabot Monarch 120 (Glucose + Hydroxyethylcellulose binder, 350°C/2h/N ₂)	8.7	9.2	2.1	15
2	Cabot Monarch 280 (Glucose + Hydroxyethylcellulose binder, 800°C/2h/N ₂)	5.4	7.7	1.7	23
3	Cabot Monarch 700 (Glucose + Hydroxyethylcellulose binder, 350°C/2h/N ₂)	12.3	12.3	2.8	31
4	Cabot Monarch 700 (Glucose + Hydroxyethylcellulose binder, 500°C/2h/N ₂)	6.3	6.3	1.4	34
5	Cabot Monarch 700 (Glucose + Hydroxyethylcellulose binder, 800°C/2h/N ₂)	18.7	18.5	4.2	31
6	Cabot Vulcan XC72 (Glucose + Hydroxyethylcellulose binder, 800°C/2h/N ₂)	5.6	5.6	1.2	30
7	Cabot Vulcan XC72R (Glucose + Hydroxyethylcellulose binder, 350°C/2h/N ₂)	7.8	7.8	1.8	38
8	Timcal Ensaco 250P (Glucose + Hydroxyethylcellulose binder, 350°C/2h/N ₂)	8.6	8.6	1.9	50
9	Timcal Ensaco 250G (Glucose binder, 800°C/4h/N ₂)	10.9	10.9	2.5	32
10	Timcal Ensaco 250G (Glucose binder, 200°C/3h/Air)	6.2	6.2	1.4	29
11	Timcal Ensaco 250G (Glucose binder, 250°C/3h/N ₂)	3.8	3.8	0.9	29

Example 5. Preparation of Catalyst Compositions

[0138] The Cabot Vulcan XC72 carbon black extrudates prepared in Example 1 were further cut into small pieces of 0.5 cm long for testing. An aqueous solution (13 ml) containing 0.17 g Au in the form of Me_4NAuO_2 and 0.26 g Pt in the form of $\text{PtO}(\text{NO}_3)$ was mixed with 21.5 g of these extrudates. The mixture was agitated to impregnate the carbon black support and was dried in a 60°C oven overnight under a dry air purge. The sample was then reduced at 350°C under forming gas (5% H_2 and 95% N_2) atmosphere for 4 hours with 2°C/min temperature ramp rate. The final catalyst was composed of 0.80 wt.% Au and 1.2 wt.% Pt.

[0139] By using other carbon black extrudates prepared from the method described above, a series of Pt-Au extrudate catalysts spanning ranges in Au and Pt loadings, Pt/Au ratios and metal distributions (e.g., eggshell, uniform, subsurface bands) could be prepared.

[0140] The cross section of a sample of the catalyst extrudate prepared with Cabot Vulcan XC72 carbon black was analyzed using scanning electron microscopy. FIG. 1 provides an image of this analysis. The image shows that platinum and gold metal was deposited on the external surface of the carbon black extrudate forming a shell coating the outer surfaces the carbon black extrudate. FIG. 2 provides a magnified view of one of catalyst extrudate cross-sections with measurements of the diameter of the carbon black extrudate (i.e., 1.14 mm) and thickness of the platinum and gold shell (average of 100 μm) on the carbon black extrudate outer surface.

Example 6. Testing of Au/Pt Carbon Black Extrudate Catalysts (using Cabot Monarch 700) in a Fixed-Bed Reactor for the Oxidation of Glucose to Glucaric Acid

[0141] Extrudates based on carbon black Cabot Monarch 700 with glucose and hydroxyethylcellulose binder and subsequent catalyst with 0.80 wt.% Au and 1.20 wt.% Pt were prepared by mixing carbon black Cabot Monarch 700 (42.0 g) and a binder solution (145.8 g prepared by heating a solution containing 3.4 wt% hydroxyethylcellulose and 28.6 wt% glucose at 80°C overnight), The resultant paste was loaded into a syringe and the material was extrudated into spaghetti-like strings with a 1.5 mm diameter. After drying in a 100°C oven for 3 hours under a dry air purge, these strings were cut into smaller pieces (2-6mm lengths) and pyrolyzed at 350°C for 2 hours under a nitrogen atmosphere. The final carbonized binder content in carbon extrudates was 31 wt%. The catalyst was then prepared using the method described in Example 5. Oxidation of glucose reactions were conducted in a 12.7 mm (0.5-inch) OD by 83 cm long 316 stainless steel tube with co-current down-flow of gas and liquid. Catalyst beds were vibration packed with 1.0 mm glass beads at the top to approximately 8 cm depth, followed by catalyst (67 cm bed depth containing 20.0 g, 0.80 wt.% Au + 1.2wt.% Pt on Cabot Monarch 700 carbon black pellets with a length of 0.5 cm and diameter of 1.5 mm prepared using the method described in Example 3), then 1.0 mm glass beads at the bottom to approximately 8 cm depth. Quartz wool plugs separated the catalyst bed from the glass beads.

[0142] The packed reactor tube was clamped in an aluminum block heater equipped with PID controller. Gas (compressed dry air) and liquid flows were regulated by mass flow controller and HPLC pump, respectively. A back pressure regulator controlled reactor pressure as indicated in Table 6. The catalyst was tested for approximately 350 hours of time on stream (TOS).

[0143] Table 6 describes the fixed bed reactor conditions and resultant extrudate catalyst performance. The catalyst productivity in Table 6 is 35 gram (glucaric acid) per gram (Pt+Au)⁻¹ hr⁻¹ or 0.70 gram (glucaric acid) per gram (catalyst)⁻¹ hr⁻¹.

Table 6.

Reactor block temperature/ °C	Glucose feed concentration/ wt. %	Reactor pressure/ psi	Liquid flowrate/ mL min ⁻¹	Gas flowrate/ mL min ⁻¹ (STP)	Glucose conversion/ %	Glucaric acid yield/ %	Selectivity/ %
130	20	750	2.00	768	>99	50	86

[0144] BET surface area measurements and BJM pore volume distribution measurements were made on the following carbon black and extrudate samples:

Sample 1: Monarch 700 carbon black material.

Sample 2: Fresh Monarch 700 extrudate prepared in accordance with this Example.

Sample 3: Monarch 700 extrudate of Example 4 following 350 hours on stream in a fixed bed reactor (described in Example 6).

Sample 4: An aqueous solution (915.0 g) containing 4.0 wt% hydroxyethylcellulose (*Sigma-Aldrich, SKU 54290, viscosity 80-125 cP (80-125 mPa. s), 2% in H₂O (20 °C)*) and 56.0 wt% glucose (*ADM Corn Processing, Dextrose Monohydrate 99.7DE with 91.2255 wt% Glucose content*) was prepared by stirring 36.6 g hydroxyethylcellulose and 561.7 g Dextrose Monohydrate in 316.7 ml D.I. water at 80°C for 16 hours. After cooling to ambient temperature, this viscous solution was added to 400.0 g carbon black powder (Cabot Monarch 700) in a blender/kneader and the material was mixed/kneaded for 1 hour. The material was then loaded into a 1" Bonnot BB Gun Extruder and extrudated into spaghetti like strings with ca. 1.5 mm diameter at cross section. These strings were dried under a dry air purge in a 120°C oven for 16 hours and then pyrolyzed at 800°C for 2 hours with 5°C/min ramp rate under a nitrogen purge. The final carbonized binder content was to be 36 wt%.

Sample 5: Prepared as described by Example 9.

Sample 6: Prepared as described by Example 12.

Sample 7: An aqueous solution (166.0 g) containing 4.0 wt% hydroxyethylcellulose (*Sigma-Aldrich, SKU54290, viscosity 80-125 cP (80-125 mPa. s), 2% in H₂O (20 °C)*) and 56.0 wt% glucose (*ADM Corn Processing, Dextrose Monohydrate 99.7DE with 91.2255 wt% Glucose content*) was prepared by stirring 6.64 g hydroxyethylcellulose and 84.8 g Dextrose Monohydrate in 74.6 ml D.I. water at 80°C for 16 hours. After cooling to ambient temperature, this viscous solution was added to 60.0 g carbon powder (Asbury 5368) in a blender/kneader and the material was mixed/kneaded for 1 hour. The material was then loaded into a 1" Bonnot BB Gun Extruder and extrudated into spaghetti like strings with ca. 1.5 mm diameter at cross section. These strings were dried under a dry air purge in a 120°C oven for 16 hours and then pyrolyzed at 800°C for 2 hours with 5°C/min ramp rate under a nitrogen purge. The final carbonized binder content was 40 wt%.

Sample 8: Commercially available activated carbon extrudate Süd Chemie G32H-N-75.

Sample 9: Commercially available activated carbon extrudate Donau Supersorbon K4-35.

[0145] The results are presented in Table 7.

Table 7.

Sample	BET Surface Area (m ² /g)	Mean Pore Diameter (nm)	BJH Pore Volume (cm ³ /g)	Pores < 3nm (% of BJH pore volume)	Pores between 10 and 50 nm (% of BJH pore volume)	Pores between 10 and 100 nm (% of BJH pore volume)	Carbonized Binder Content (wt %)
Sample 1	180	12	0.38	5	40	75	0
Sample 2	178	11	0.29	7	45	75	36
Sample 3	98	18	0.31	3	50	90	36
Sample 4	182	13	0.36	4	55	80	36
Sample 5	194	11	0.29	6	45	75	36
Sample 5	234	10	0.29	7	45	70	36
Sample 6	218	12	0.33	5	60	80	40
Sample 8	1164	3.4	0.63	40	<15	15	**
Sample 9	1019	2.7	0.31	65	5	7	**

[0146] FIG. 3 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a raw Monarch 700 carbon black material. FIG. 4 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a fresh catalyst prepared from a carbon black extrudate using Monarch 700 and a glucose/hydroxyethylcellulose binder. FIG. 5 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for the catalyst extrudate of FIG. 2 following 350 hours of use in a fixed bed reactor for the oxidation of glucose to glucaric acid. FIG. 6 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a extrudate using Monarch 700 carbon black and a glucose/hydroxyethyl cellulose binder. FIG. 7 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a extrudate using Sid Richardson SC 159 carbon black and a glucose/hydroxyethyl cellulose binder. FIG. 8 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a extrudate using Sid Richardson SC 159 carbon black and a glucose/hydroxyethyl cellulose binder prepared in accordance with Example 12. FIG. 9 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a extrudate using Asbury 5368 carbon black and a glucose/hydroxyethyl cellulose binder. FIG. 10 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a commercially available activated carbon extrudate of Süd Chemie G32H-N-75. FIG. 11 presents a plot of the cumulative pore volume (%) as a function of mean pore diameter for a commercially available activated carbon extrudate of Donau Supersorbon K4-35.

[0147] FIG. 12 presents the pore size distribution for an extrudate using Sid Richardson SC159 carbon black and a glucose/hydroxyethyl cellulose binder measured by mercury porosimetry. These plots show that the micropore contribution to pore volume for carbon black extrudate catalysts (fresh and after use) is

very low. In particular the plots show that the micropore contribution (pores < 3nm) is less than 10% of the BJH pore volume. In some instances the micropore contribution (pores < 3nm) is less than 6% of the BJH pore volume, and in some instances the micropore contribution (pores < 3nm) is less than 4% of the BJH pore volume. In contrast, the micropore contribution to pore volume for an activated carbon extrudate catalyst is exceedingly high at 40%. Also, the plots show that the contribution to pore volume from pores having a mean diameter from 10 nm to 50 nm for the carbon black catalysts was 40% or higher. On the other hand, the contribution to pore volume from pores having a mean diameter from 10 nm to 50 nm for the activated carbon catalyst was less than 15%. The plots show that the contribution to pore volume from pores having a mean diameter from 10 nm to 100 nm for the carbon black catalysts was 70% or higher. On the other hand, the contribution to pore volume from pores having a mean diameter from 10 nm to 100 nm for the activated carbon catalyst was 15% or less.

Example 7. Testing of Au/Pt Carbon Black Extrudate Catalysts (using Cabot Vulcan XC72) in a Fixed-Bed Reactor for the Oxidation of Glucose to Glucaric Acid

[0148] Extrudates based on carbon black Cabot Vulcan XC72 and subsequent catalyst with 0.80 wt.% Au and 1.20 wt.% Pt were prepared by mixing carbon black Cabot Vulcan XC72 (36.4 g) and a binder solution (136.5 g prepared by heating a solution containing 3.7 wt% hydroxyethylcellulose and 24.4 wt% glucose at 80°C overnight). The resultant paste was loaded into a syringe and the material was extrudated into spaghetti-like strings with a 1.5 mm diameter followed by drying at 120°C for 4 hours in air, and pyrolysis at 350°C for 2 hours under a nitrogen atmosphere. The final binder content in pyrolyzed carbon extrudates was 30 wt%. The catalysts were prepared using the method described in Example 3. The catalyst was tested in the same 12.7 mm (0.5-inch) OD fixed-bed reactor as in Example 6. Table 8 describes the fixed bed reactor conditions and resultant extrudate catalyst performance. The catalyst productivity in Table 8 is 36 gram (glucaric acid) per gram (Pt+Au)⁻¹hr⁻¹ or 0.72 gram (glucaric acid) per gram (catalyst)⁻¹ hr⁻¹.

Table 8.

Reactor block temperature/ °C	Glucose feed concentration/ wt. %	Reactor pressure/ psi	Liquid flowrate/ mL min ⁻¹	Gas flowrate/ mL min ⁻¹ (STP)	Glucose conversion/ %	Glucaric acid yield/ %	Selectivity /%
130	20	750	2.00	768	>99	52	87

Example 8. Oxidation of Glucose to Glucaric Acid - Activated Carbons with High Surface Areas (Comparative)

[0149] The same synthesis procedure described in Example 6 was used to prepare Pt-Au catalysts supported on high surface area activated carbon. The activated carbon extrudates were crushed and sieved to < 90 μm prior to the catalyst preparation and screening. The catalysts were screened in the same reactor under the same conditions described in Example 2(B)(ii). As shown in Table 9, the high surface area activated carbon carriers were found to exhibit lower activity and lower selectivity (as defined herein).

Table 9.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Glucaric Acid Yield (%)	Selectivity (%)
Donau Supersorbon K4-35	1019	27	0.31	22	66
Donau Supersorbon SX30	1050	39	0.71	18	68
Norit RX3 Extra	1239	37	0.23	20	70

Example 9. Preparation of Carbon Black Extrudate Catalysts - Attrition and Abrasion Testing

[0150] An aqueous solution (113 g) containing 4.0 wt% hydroxyethylcellulose (Sigma-Aldrich, SKU 54290, viscosity 80-125 cP (80-125 mPa. s), 2% in H₂O (20°C)) and 56.0 wt% glucose (ADM Corn Processing, Dextrose Monohydrate 99.7DE with 91.2255 wt% Glucose content) was prepared by stirring 4.5 g hydroxyethylcellulose and 69.4 g Dextrose Monohydrate in 39.1 ml D.I. water at 80°C overnight. After cooling to room temperature, this viscous solution was added to 50 g carbon black powder (Sid Richardson SC159, 231 m²/g) in a blender/kneader and the material was mixed/kneaded for 1 hour. The material was then loaded into a 1" Bonnot BB Gun Extruder and extruded into spaghetti like strings with ca. 1.5 mm diameter at cross section. These strings were dried under a dry air purge in a 120°C oven overnight and then pyrolyzed at 800°C for 4 hours with 5°C/min ramp rate under a nitrogen purge. The extruded and pyrolyzed samples were cut into small pieces of 0.5 cm long for testing.

[0151] The properties of the resultant extrudate are show in Table 10. BET and crush strength measurements performed as described in the current disclosure.

Table 10.

BET Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Single Piece Crush Strength (N)	Mean Radial Piece Crush Strength (N/mm)
194	112	0.29	90	30

[0152] Extrudates prepared in accordance with this example were tested for the determination of attrition index (ASTM D4058-96) and abrasion loss according to the procedure described below.

Measurement of Attrition Index.

[0153] The ASTM attrition index (ATTR) is a measurement of the resistance of a catalyst or extrudate particle to attrition wear, due to the repeated striking of the particle against hard surfaces within the specified test drum. The diameter and length of the drum is similar to that described in ASTM D4058, with a rolling apparatus capable of delivering 55 to 65 RPM of rotation to the test drum. The percentage of the original sample that remains on a 20-mesh sieve is called the "Percent Retained" result of the test. The results of the test can be used, on a relative basis, as a measure of fines production during the handling, transportation, and use of the catalyst or extrudate material. A percent retained result of > 97% is desirable for an industrial application.

[0154] Approximately 100 g of the extrudate material prepared in Example 9 above was transferred to the test drum which was fastened and transferred to the rolling apparatus and rolled at 55 to 65 RPM for 35 minutes. The weight percent retained after the test was 99.7 %.

Measurement of Abrasion Loss.

[0155] The abrasion loss (ABL) is an alternate measurement of the resistance of a catalyst or extrudate particle to wear, due to the intense horizontal agitation of the particles within the confines of a 30-mesh sieve. The results of this procedure can be used, on a relative basis, as a measure of fines production during the handling, transportation, and use of the catalyst or adsorbent material. An abrasion loss of < 2 wt% is desired for an industrial application. Approximately 100 g of the extrudate material prepared in example 9 above was first de-dusted on a 20-mesh sieve by gently moving the sieve side-to-side at least 20 times. The de-dusted sample was then transferred to the inside of a clean, 30-mesh sieve stacked above a clean sieve pan for the collection of fines. The complete sieve stack was then assembled onto a RO-Tap RX-29 sieve shaker, covered securely and shaken for 30 minutes. The fines generated were weighed to provide a sample abrasion loss of 0.016 wt.%.

Example 10. Testing of Au/Pt Carbon Black Extrudate Catalysts of Example 9 in a Fixed-Bed Reactor for the Oxidation of Glucose to Glucaric Acid.

[0156] The carbon black extrudates prepared from the method described in Example 9 were further cut into small pieces of 0.5 cm long for testing. To 27.0 g of these extrudates, an aqueous solution (8.0 ml) containing 0.16 g Au in the form of Me_4NAuO_2 and 0.24 g Pt in the form of $\text{PtO}(\text{NO}_3)$ was added. The mixture was agitated to impregnate the carbon black support and was dried in a 70°C oven for 1 hour under a dry air purge. The sample was then reduced at 350°C under forming gas (5% H_2 and 95% N_2) atmosphere for 4 hours with 2°C/min temperature ramp rate. The final catalyst was composed of ca. 0.60 wt.% Au and 0.90 wt.% Pt. By using other carbon black extrudates prepared from the methods described herein, a series of Pt-Au extrudate catalysts spanning ranges in Au and Pt loadings, Pt/Au ratios and metal distributions (e.g., eggshell, uniform, subsurface bands) can be prepared.

[0157] The glucose to glucaric acid oxidation reaction was conducted in a ½" OD by 83 cm long 316 stainless steel tube with co-current down-flow of gas and liquid. Catalyst beds were vibration packed with 1.0 mm glass beads at the top to approximately 10 cm depth, followed by catalyst (63 cm bed depth containing 27.4 g, 0.60wt% Au + 0.90wt% Pt on Sid Richardson SC159 carbon black pellets with a length of 0.5 cm and diameter of 1.4 mm prepared using the method described, then 1.0 mm glass beads at the bottom to approximately 10 cm depth. Quartz wool plugs separated the catalyst bed from the glass beads.

[0158] The packed reactor tube was clamped in an aluminum block heater equipped with PID controller. Gas (compressed dry air) and liquid flows were regulated by mass flow controller and HPLC pump, respectively. A back pressure regulator controlled reactor pressure as indicated in Table 11. The catalyst was tested for ca. 920 hours on stream and showed stable performance. Table 11 describes the fixed bed reactor conditions and resultant extrudate catalyst performance. The catalyst productivity in Table 11 is 23 gram (glucaric acid) per gram $(\text{Pt}+\text{Au})^{-1}\text{hr}^{-1}$ or 0.35 gram (glucaric acid) per gram $(\text{catalyst})^{-1}\text{hr}^{-1}$.

Table 11.

Reactor block temperature/ °C	Glucose feed concentration/ wt%	Reactor pressure/ psi	Liquid flowrate/ mL min ⁻¹	Gas flowrate/ mL min ⁻¹ (STP)	Glucose conversion/ %	Glucaric acid yield/%	Selectivity/
125	20	750	2.00	512	>99	32	79

[0159] After 920 hours on stream the catalyst extrudate was removed and re-submitted for mechanical crush strength testing. The mean single piece crush strength and mean radial piece crush strength data were found be within experimental error unchanged from the data listed in Table 10, thereby illustrating that the extrudate catalyst prepared by the method described is productive, selective and stable under the continuous flow conditions described.

Example 11. Testing of Au/Pt Carbon Black Extrudate Catalysts (using Asbury 5368) in a Fixed-Bed Reactor for the Oxidation of Glucose to Glucaric Acid

[0160] Reactions were conducted in a ½" OD by 83 cm long 316 stainless steel tube with co-current down-flow of gas and liquid. Catalyst beds were vibration packed with 1.0 mm glass beads at the top to approximately 8 cm depth, followed by catalyst (73 cm bed depth containing 35.0 g, 0.50 wt% Au + 0.85 wt% Pt on Asbury 5368 extruded pellets (as previously described for Sample 7 (Example 6) with a length of 0.5 cm and diameter of 1.4 mm prepared using the method described in previous example 9), then 1.0 mm glass beads at the bottom to approximately 8 cm depth. Quartz wool plugs separated the catalyst bed from the glass beads.

[0161] The packed reactor tube was clamped in an aluminum block heater equipped with PID controller. Gas (compressed dry air) and liquid flows were regulated by mass flow controller and HPLC pump, respectively. A back pressure regulator controlled reactor pressure as indicated in Table 12. The catalyst was tested for ca. 240 hours TOS and showed stable performance. Table 12 describes the fixed bed reactor conditions and resultant extrudate catalyst performance. The catalyst productivity in Table 12 is 20 gram (glucaric acid) per gram (Pt+Au)⁻¹hr⁻¹ or 0.27 gram (glucaric acid) per gram (catalyst)⁻¹ hr⁻¹.

Table 12. 0.50 wt% Au + 0.85 wt% Pt / Asbury 5368 Extrudate (stable performance over 240 hours on stream)

Reactor block temperature/ °C	Glucose feed concentration/ wt%	Reactor pressure/ psi	Liquid flowrate/ mL min ⁻¹	Gas flowrate/ mL min ⁻¹ (STP)	Glucose conversion/ %	Glucaric acid yield/%	Selectivity
125	20	750	2.00	512	>99	31	76

Example 12. Preparation of Carbon Black Extrudate Catalysts on Partially Oxidized Support

[0162] Sid Richardson SC159 carbon black extrudates prepared from the method described in example 9 were oxidized in air at 300°C for 3 hours with 5°C/min ramp rate to give partially oxidized pellets. To 36.0 g of these partially oxidized extrudates, an aqueous solution (9.0 ml) containing 0.18 g Au in the form of Me₄NAuO₂ and 0.31 g Pt in the form of PtO(NO₃) was added. The mixture was agitated to impregnate the carbon black support and was dried in a 60°C oven overnight under a dry air purge. The sample was then

reduced at 350°C under forming gas (5% H₂ and 95% N₂) atmosphere for 4 hours with 2°C/min temperature ramp rate. The final catalyst was composed of ca. 0.50 wt% Au and 0.85 wt% Pt. By using other carbon black extrudates prepared from the method described herein, a series of Pt-Au extrudate catalysts spanning ranges in Au and Pt loadings, Pt/Au ratios and metal distributions (e.g., eggshell, uniform, subsurface bands) could be prepared. The glucose to glucaric acid oxidation reaction was conducted in a ½" OD by 83 cm long 316 stainless steel tube with co-current down-flow of gas and liquid. Catalyst beds were vibration packed with 1.0 mm glass beads at the top to approximately 6 cm depth, followed by catalyst (70.4 cm bed depth containing 34.5 g, 0.50wt% Au + 0.85wt% Pt on partially oxidized Sid Richardson SC159 carbon black pellets with a length of 0.5 cm and diameter of 1.5 mm prepared using the method described in Example 2), then 1.0 mm glass beads at the bottom to approximately 6 cm depth. Quartz wool plugs separated the catalyst bed from the glass beads.

[0163] The packed reactor tube was clamped in an aluminum block heater equipped with PID controller. Gas (compressed dry air) and liquid flows were regulated by mass flow controller and HPLC pump, respectively. A back pressure regulator controlled reactor pressure as indicated in Table 13. The catalyst was tested for ca. 230 hours TOS and showed stable performance. Table 13 describes the fixed bed reactor conditions and resultant extrudate catalyst performance. The catalyst productivity in Table 13 is 26 gram (glucaric acid) per gram (Pt+Au)⁻¹hr⁻¹ or 0.36 gram (glucaric acid) per gram (catalyst)⁻¹ hr⁻¹.

Table 13.

Reactor block temperature/ °C	Glucose feed concentration/ wt%	Reactor pressure/ psi	Liquid flowrate/ mL min ⁻¹	Gas flowrate/ mL min ⁻¹ (STP)	Glucose conversion/ %	Glucaric acid yield/ %
125	20	750	2.00	512	>99	42

Example 13. Preparation of Carbon Black Extrudates using a Poly(vinylalcohol) Porogen

[0164] An aqueous solution (490.0 g) containing 8.0 wt% Mowiol 8-88 Poly(vinylalcohol) (Mw 67k, Sigma-Aldrich 81383) and 36.0 wt% glucose (ADM Corn Processing, Dextrose Monohydrate 99.7DE with 91.2255 wt% Glucose content) was prepared by stirring 39.2 g Mowiol 8-88 Poly(vinylalcohol) and 193.4 g Dextrose Monohydrate in 257.4 ml D.I. water at 70°C overnight. After cooling to room temperature, this solution was added to 230 g carbon black powder (Sid Richardson SC159) in a blender/kneader and the material was mixed/kneaded for 1 hour. The material was then loaded into a 1" Bonnot BB Gun Extruder and extruded into spaghetti like strings with ca. 1.5 mm diameter at cross section. These strings were further dried in a 90°C oven overnight under a dry air purge and then pyrolyzed at 600°C for 4 hours with 5°C/min ramp rate in a nitrogen atmosphere. The final carbonized binder content was 24 wt.%. The resultant extrudate (3-5 mm in length) possessed a surface area of 149 m²/g, a pore volume of 0.35 cm³/g and a mean pore diameter of 16 nm. The mean radial piece crush strength of these pellets was measured to be 11.5 N/mm. The single piece crush strength was measured to be 42N.

Example 14. Testing of Au/Pt Activated Carbon Extrudate Catalysts (using Clariant Donau Supersorbon K4-35 Activated Carbon Extrudate) in a Fixed-Bed Reactor for the Oxidation of Glucose to Glucaric Acid

[0165] Catalyst based on activated carbon Clariant Supersorbon K 4-35 was prepared using the same

method described in Example 7. Oxidation of glucose reactions were conducted using the same method described in Example 2(B)(ii). A catalyst bed depth of 73cm containing 27.0 g, 0.53 wt.% Au + 0.90 wt.% Pt on Clariant Supersorbon K 4-35 activated carbon pellets with a length of 0.5 cm and diameter of 1.4 mm was tested for approximately 40 hours of time on stream (TOS). Table 14 describes the fixed bed reactor conditions and resultant extrudate catalyst performance. After 40 hours on stream the glucaric acid yield and the catalyst productivity were determined to be lower than the shaped carbon black catalysts of the invention.

Table 14.

Reactor block temperature/ °C	Glucose feed concentration/ wt%	Reactor pressure/ psi	Liquid flowrate/ mL min ⁻¹	Gas flowrate/ mL min ⁻¹ (STP)
125	20	750	2.00	512

Example 15. Preparation of Carbon Black Extrudates

[0166] An aqueous solution (915 g) containing 4.0 wt% hydroxyethylcellulose (HEC) (Sigma-Aldrich, SKU 54290, viscosity 80-125 cP (80-125 mPa. s) at 2% in H₂O (20°C)) and 56.0 wt% glucose (ADM Corn Processing, Dextrose Monohydrate 99.7DE with 91.2 wt% Glucose content) was prepared by stirring 36.6 g hydroxyethylcellulose and 561.7 g Dextrose Monohydrate in 316.7 ml D.I. water at approximately 80°C for 2 hours. To this viscous solution was added 400.1 g of Sid Richardson SC159 carbon black powder, the mixture was then mixed for a further 10 minutes. The material was then loaded into a 1" diameter Bonnot extruder, fitted with a ¼ inch spacer and a die with 1.6 mm cylindrical holes, and extruded into spaghetti-like strings. The extrudate was dried in a 110°C oven overnight, then pyrolyzed in a stationary lab furnace under a nitrogen purge at 800°C for 4 hours (after ramping the temperature up at 5°C/minute to reach the target temperature) (Table 15).

Table 15. Properties of the pyrolyzed extrudate from Example 15

N ₂ BET Surface Area (m ² /g)	BJH N ₂ Pore Volume (cm ³ /g)	Diameter at cross section (mm)	Radial Piece Crush Strength (N/mm)
207	0.30	1.5	17

Example 16. Preparation of Carbon Black Extrudates

[0167] An aqueous solution (3813 g) containing 4.0 wt% hydroxyethylcellulose (HEC) (Sigma-Aldrich, SKU 54290, viscosity 80-125 cP (80-125 mPa. s) at 2% in H₂O (20°C)) and 56.0 wt% glucose (ADM Corn Processing, Dextrose Monohydrate 99.7DE with 91.2 wt% Glucose content) was prepared by stirring 153 g hydroxyethylcellulose and 2340 g Dextrose Monohydrate in 1320 ml D.I. water at approximately 80°C for 3 hours. This viscous solution was added over 3.5 minutes to 1670 g of Sid Richardson SC159 carbon black powder in a mix-muller, the mixture was then mixed for a further 20 minutes in the mix-muller. The material was then loaded into a 2" diameter Bonnot extruder, fitted with 5 dies with 26 cylindrical holes each 1/16" internal diameter (JMP Industries, part number 0388P062), and no spacer, and extruded into spaghetti-like strings. 1515 g of the extrudate was dried in a 110°C oven overnight, to produce 1240g of dried extrudate. The product was then pyrolyzed in a stationary tube furnace under a nitrogen purge at 800°C for 4 hours (Table 16).

Table 16. Properties of the pyrolyzed extrudate from Example 2

N ₂ BET Surface Area (m ² /g)	BJH N ₂ Pore Volume (cm ³ /g)	Diameter at cross section (mm)	Radial Piece Crush Strength (N/mm)
169	0.23	1.4	13

Example 17. Preparation of Carbon Black Extrudates, using Batch Pyrolysis in a Rotary Tube Furnace

[0168] An aqueous solution (3813 g) containing 4.0 wt% of Dow Cellosize HEC QP 40 hydroxyethylcellulose (viscosity 80-125 cP (80-125 mPa. s) at 2% in H₂O (20°C)), and 56.0 wt% glucose (ADM Corn Processing, Dextrose Monohydrate 99.7DE with 91.2 wt% Glucose content) was prepared by stirring 153 g hydroxyethylcellulose and 2340 g Dextrose Monohydrate in 1320 ml D.I. water at approximately 80°C for 3 hours. This viscous solution was added over 3.5 minutes to 1670 g of Sid Richardson SC159 carbon black powder in a mix-muller, the mixture was then mixed for a further 20 minutes in the mix-muller. The material was then loaded into a 2" diameter Bonnot extruder, fitted with 5 dies with 26 cylindrical holes each 1/16" internal diameter (JMP Industries, part number 0388P062), and no spacer, and extruded into spaghetti-like strings. 3.9 kg of the extrudate was dried in a 110°C oven overnight, to produce 2.93 kg of dried extrudate. This dried extrudate was then screened over an 18 mesh screen, and 2.91 kg of screened material were collected.

[0169] The mixing, extrusion, drying and screening procedure above was repeated 3 more times to generate a total of 4 batches of dried, screened extrudate, which were combined, as summarized in Table 17.

Table 17. Production of Carbon Black Extrudate in a 2" Extruder

Sample number	Mass of wet extrudate collected (kg)	Mass of dried extrudate (kg)	Mass of dried extrudate after screening (kg)
17.1	3.90	2.93	2.91
17.2	4.66	3.80	3.76
17.3	5.17	4.24	4.20
17.4	4.85	3.74	3.71
Combined Total	18.58	14.71	14.58

[0170] 650g batches of the combined dried & screened extrudate of were then pyrolyzed in a rotary tube furnace under a nitrogen purge at 800°C for 2 hours, each batch producing approximately 350g of pyrolyzed product. For each batch, 650g of the carbon black extrudates (prepared from Sid Richardson SC159 with glucose and hydroxyethylcellulose binders) were loaded into an MTI Corporation 5" Quartz Tube Three Zone Rotary Tube Furnace (OTF-1200X-5L-R-III-UL). The carbon black extrudates were pyrolyzed with the 5" quartz tube rotating at 4.0 rpm under a nitrogen atmosphere at 800°C for 2 hours with the following temperature ramp: ambient temperature to 200°C at 10°C/min, 200°C to 600°C at 5°C/min, 600°C to 800°C at 10°C/min, hold at 800°C for 2 hours, then allowed to cool to ambient temperature, still under nitrogen purge. 350 g of pyrolyzed carbon black extrudates were recovered, with 51.5% yield by mass. The properties of the batch-pyrolyzed extrudate are shown in Table 18. Other carbon black extrudates can be pyrolyzed at various temperatures in a similar manner, or using a continuously operating rotary kiln as described in the next example.

Table 18. Properties of Carbon Black Extrudate Batch-Pyrolyzed in a Rotary Tube Furnace

	N ₂ BET Surface Area (m ² /g)	N ₂ Mean Pore Diameter (Å)	BJH N ₂ Pore Volume (cm ³ /g)	Radial Piece Crush Strength (N/mm)
Combined Extrudates of Example 17 Batch- Pyrolyzed at 800°C for 2 hours	191	100	0.29	15

Example 18. Preparation of Carbon Black Extrudates, using Continuous Pyrolysis in Rotary Kiln

[0171] The mixing, extrusion, drying and screening procedure described in Example 17 was repeated 10 more times to generate an additional 33.4 kg of dried, screened extrudate, which were combined. 25.7 kg of the combined dried & screened extrudate was then pyrolyzed in a continuous rotary kiln, with a continuous nitrogen purge (counter current flow vs. the extrudate), with a continuous feed of dried extrudate at approximately 0.5 kg/hour, with product collected at a number of set point conditions summarized in Table 19. The rotary kiln was electrically heated; the temperature set points for the external heaters are shown in Table 19, along with the calculated residence time of the material in the heating zone. The temperature and residence time were adjusted to influence the surface area of the product. A total of 12.5 kg of pyrolyzed product was collected, for an overall mass-based yield of 48.5%.

Table 19. Properties of Carbon Black Extrudate Pyrolyzed in a Continuously Operated Rotary Kiln

Continuous Rotary Kiln Pyrolyzed Extrudate Sample Number	Temperature Set Point in Heating Zone (°C)	Calculated Residence Time in Heating Zone (minutes)	N ₂ Specific Surface Area (m ² /g)
18.1	820	63	219
18.2	820	44	208
18.3	800	44	202
18.4	780	44	188
18.5	760	44	188
18.6	780	33	176
18.7	820	33	181

Example 19. Hydrodeoxygenation of Glucaric Acid Dilactone to Adipic Acid

[0172] Suitably concentrated aqueous solutions of rhodium nitrate and platinum nitrate were added together to carbon black powder (crushed from carbon black pellets) by incipient wetness impregnation and agitated to impregnate supports. The samples were dried in an oven at 60°C overnight, and reduced at 350°C under a forming gas (5% H₂ and 95% N₂) atmosphere for 4 hours with 2°C/min temperature ramp rate to produce catalysts with a composition of 1.0 wt.% Rh and 2.0 wt.% Pt. By using other carbon blacks supports, Rh and Pt precursors, and adjusting amount of Rh and Pt in solution, different catalysts with various Rh and Pt loadings on a variety of particles from extrudates were prepared in a similar manner.

[0173] These catalysts were tested for hydrodeoxygenation of glucaric acid dilactone using the following

testing protocol. Catalyst (16 mg) was weighed into a glass vial insert followed by addition of a solution (125 μ l) containing glucaric acid dilactone (0.80 M), HBr (0.80 M), and water (2.0 M). The glass vial insert was loaded into a reactor and the reactor was closed. The atmosphere in the reactor was replaced with hydrogen and pressurized to 900 psig (6306 kPa) at room temperature. Reactor was heated to 120°C and maintained at 120°C for 1 hour while vials were shaken. Reactor was then heated to 160°C and maintained at 160°C for 2 hours while vials were shaken. After that, shaking was stopped and reactor was cooled to 40°C. Pressure in the reactor was slowly released. The glass vial insert was removed from reactor and centrifuged. The clear solution was hydrolyzed with NaOH, diluted with deionized water, and analyzed by ion chromatography to determine the yield of adipic acid. The properties of the carbon black starting materials and results of the reaction screening are presented in Table 20.

Table 20.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Adipic Acid Yield (%)	Adipic Acid Selectivity (%)
Cabot Monarch 120	25	277	0.1	75	83
Cabot Monarch 280	30	176	0.1	81	91
Timcal Ensaco 250P	64	140	0.24	92	99
Cabot Monarch 570	102	138	0.3	87	97
Cabot Monarch 700	181	121	0.38	89	99
Cabot Vulcan XC72	224	161	0.43	86	99
Sid Richardson SC159	234	182	0.81	75	86

Example 20. Testing of Rh/Pt Carbon Black Extrudate Catalysts in a Fixed-Bed Reactor for Hydrodeoxygenation of Glucaric Acid to Adipic Acid

[0174] Cabot Vulcan XC72 carbon black particles used in this experiment were 150 to 300 μ m sized particles crushed and sieved from extrudate pellets prepared from the method described in previous examples. Reactions were conducted in a 6.4 mm (0.25 inch) OD by 38 cm long zirconium tube with co-current down-flow of gas and liquid. Catalyst beds were vibration packed with 200 to 300 μ m sized glass beads at the top to approximately 5 cm depth, followed by catalyst (28 cm bed depth containing 1.9 g, 0.90 wt.% Rh + 2.1 wt.% Pt on carbon black particles, 150 to 300 μ m particle size), then 200 to 300 μ m sized glass beads at the bottom to approximately 5 cm depth. Quartz wool plugs separated the catalyst bed from the glass beads.

[0175] The packed reactor tube was clamped in an aluminum block heater equipped with PID controller. Gas (compressed hydrogen) and liquid flows were regulated by mass flow controller and HPLC pump, respectively. Substrate solution contains 0.80M D-glucaric acid-1,4:6,3-dilactone, 0.40M HBr and 2.0M water in acetic acid. A back pressure regulator controlled reactor pressure as indicated in Table 21. External temperature of top half reactor and bottom half reactor was controlled at 110°C and 160°C respectively. The catalyst was tested for 350 hours on stream and showed stable performance. Table 21

describes the fixed bed reactor conditions and resultant catalyst performance.

Table 21.

Test	Reactor block temperature / °C	Glucaric acid dilactone concentration /	Reactor pressure / psi	Liquid flowrate / mL min ⁻¹	Gas flowrate / mL min ⁻¹ (STP)	Glucaric acid dilactone conversion / %	Adipic acid yield / %
1	110/160	0.80	1000	0.050	50	90	42

Example 21. Hydrodeoxygenation of 1,2,6-Hexanetriol to 1,6-Hexanediol

[0176] Suitably concentrated aqueous solutions of $\text{Pt}(\text{NO}_3)_x$ and $\text{H}_4\text{SiO}_4 \cdot 12\text{WO}_3$ or PtONO_3 and $\text{H}_4\text{SiO}_4 \cdot 12\text{WO}_3$ were added to approximately 50 mg of Ensaco 250G carbon and agitated to impregnate the supports. The samples were dried in an oven at 40°C overnight under static air and then reduced at 350°C under forming gas (5% H_2 and 95% N_2) atmosphere for 3 hours. The final catalysts had a metal content of approximately 4.09 wt.% Pt and 3.42 wt.% W.

[0177] These catalysts were tested for 1,2,6-hexanetriol hydrodeoxygenation using the following catalyst testing protocol. Catalyst (approximately 10 mg) was weighed into a glass vial insert followed by addition of an aqueous 1,2,6-hexanetriol solution (200 μl of 0.8 M). The glass vial insert was loaded into a reactor and the reactor was closed. The atmosphere in the reactor was replaced with hydrogen and pressurized to 670 psig (4720 kPa) at room temperature. The reactor was heated to 160°C and maintained at the respective temperature for 150 minutes while vials were shaken. After 150 minutes, shaking was stopped and reactor was cooled to 40°C. Pressure in the reactor was then slowly released. The glass vial insert was removed from the reactor and centrifuged. The solution was diluted with methanol and analyzed by gas chromatography with flame ionization detection. The results are shown in Table 22.

Table 22.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Pt Precursor	W Precursor	Yield (%)	Selectivity (%)
Ensaco 250G	64	140	0.24	$\text{Pt}(\text{NO}_3)_2$	$\text{H}_4\text{SiO}_4 \cdot 12\text{WO}_3$	27	64
Ensaco 250G	64	140	0.24	PtONO_3	$\text{H}_4\text{SiO}_4 \cdot 12\text{WO}_3$	47	69

Example 22. Hydrodeoxygenation of 1,2,6-Hexanetriol to 1,6-Hexanediol

[0178] A suitably concentrated aqueous solution of ammonium metatungstate, $\text{H}_{26}\text{N}_6\text{W}_{12}\text{O}_{40}$ was added to approximately 500 mg of Ensaco 250G and agitated to impregnate the carbon black support. The sample was thermally treated at 600°C under a nitrogen atmosphere for 3 hours with 5 °C/min temperature ramp rate. Suitably concentrated aqueous solutions of $\text{Pt}(\text{NMe}_4)_2(\text{OH})_6$ was added to 50 mg of the above sample and agitated to impregnate the carbon supports. The samples were dried in an oven at 40°C overnight under static air and then reduced at 250°C under forming gas (5% H_2 and 95% N_2) atmosphere for 3 hours with 5 °C/min temperature ramp rate. The final catalysts had a metal content of approximately 4.5 wt.% Pt

and 2 wt.% W.

[0179] These catalysts were tested for 1,2,6-hexanetriol hydrodeoxygenation using the following catalyst testing protocol. Catalyst (approximately 10 mg) was weighed into a glass vial insert followed by addition of an aqueous 1,2,6-hexanetriol solution (200 μ l of 0.8 M). The glass vial insert was loaded into a reactor and the reactor was closed. The atmosphere in the reactor was replaced with hydrogen and pressurized to 670 psig (4720 kPa) at room temperature. The reactor was heated to 160°C and maintained at the respective temperature for 150 minutes while vials were shaken. After 150 minutes, shaking was stopped and reactor was cooled to 40°C. Pressure in the reactor was then slowly released. The glass vial insert was removed from the reactor and centrifuged. The clear solution was diluted with methanol and analyzed by gas chromatography with flame ionization detection. The results are shown in Table 23.

Table 23.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Pt Precursor	W Precursor	Yield (%)	Selectivity (%)
Ensaco 250G	64	140	0.24	Pt(NMe ₄) ₂ (OH) ₆	H ₂₆ N ₆ W ₁₂ O ₄₀	38	88

Example 23. Hydrodeoxygenation of 1,2,6-Hexanetriol to 1,6-Hexanediol

[0180] Suitably concentrated aqueous solutions of ammonium metatungstate, H₂₆N₆W₁₂O₄₀ were added to approximately 500 mg of carbon black materials and agitated to impregnate the carbon black supports. The samples were thermally treated at 600°C under a nitrogen atmosphere for 3 hours with 5 °C/min temperature ramp rate. Suitably concentrated aqueous solutions of Pt(NMe₄)₂(OH)₆ were added to approximately 50 mg of the above samples and agitated to impregnate the carbon supports. The samples were dried in an oven at 60°C under static air and then reduced at 350°C under forming gas (5% H₂ and 95% N₂) atmosphere for 3 hours with 5 °C/min temperature ramp rate. The final catalysts had a metal content of approximately 5.7 wt.% Pt and 1.8 wt.% W.

[0181] These catalysts were tested for 1,2,6-hexanetriol hydrodeoxygenation using the following catalyst testing protocol. Catalyst (approximately 10 mg) was weighed into a glass vial insert followed by addition of an aqueous 1,2,6-hexanetriol solution (200 μ l of 0.8 M). The glass vial insert was loaded into a reactor and the reactor was closed. The atmosphere in the reactor was replaced with hydrogen and pressurized to 670 psig (4720 kPa) at room temperature. The reactor was heated to 160°C and maintained at the respective temperature for 150 minutes while vials were shaken. After 150 minutes, shaking was stopped and reactor was cooled to 40°C. Pressure in the reactor was then slowly released. The glass vial insert was removed from the reactor and centrifuged. The clear solution was diluted with methanol and analyzed by gas chromatography with flame ionization detection. Results are shown in Table 24.

Table 24.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Pt Precursor	W Precursor	Yield (%)	Selectivity (%)
Ensaco 250G	64	140	0.24	Pt(NMe ₄) ₂ (OH) ₆	H ₂₆ N ₆ W ₁₂ O ₄₀	52	71

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Pt Precursor	W Precursor	Yield (%)	Selectivity (%)
Orion HiBlack 40B2	109	155	0.32	Pt(NMe ₄) ₂ (OH) ₆	H ₂₆ N ₆ W ₁₂ O ₄₀	28	68

Example 24. Small Scale Batch Reactor Experiments for Amination of 1,6-Hexanediol to produce 1,6-Hexamethylenediamine

Amination of 1,6-Hexanediol to produce 1,6-Hexamethylenediamine - Analytical Details

[0182] Product composition was determined by HPLC analysis using a Thermo Ultimate 3000 dual analytical chromatography system. Hexamethylenediamine (HMDA), hexamethyleneimine (HMI) and pentylamine were eluted with a mobile phase consisting of H₂O/MeCN/TFA and detected using a charged aerosol detector (CAD). 1,6-Hexanediol (HDO) was eluted with a mobile phase consisting of H₂O/MeCN/TFA and detected using a refractive index detector (RI). In certain examples an internal standard, N-methyl-2-pyrrolidone (NMP), was used in the substrate feed to correct for variations in product effluent concentration due to NH₃ off-gassing. NMP was eluted with a mobile phase consisting of H₂O/MeCN/TFA and detected by UV at 210nm. All products were quantified by comparison to calibration standards. Selectivity is reported as the yield of HMDA divided by the sum of HMDA and pentylamine.

Experiment 1

Preparation of Supported Ru Catalysts

[0183] A suitably concentrated aqueous solution of Ru(NO)(NO₃)₃ was added to a 96 vial array of carbon supports containing 10 or 20 mg of support in each vial. The volume of ruthenium solution was matched to equal the pore volume of the support. Each sample was agitated to impregnate the support. The samples were dried in an oven at 60°C for 12 hours under a dry air purge. The catalysts were reduced under forming gas (5% H₂ and 95% N₂) at 250 °C for 3 hours using a 2°C/min temperature ramp rate. The final catalysts were composed of 2 weight percent ruthenium.

Catalyst Screening Procedure

[0184] A substrate solution consisting of 0.7M 1,6-hexanediol in concentrated aqueous NH₄OH was added to an array of catalysts prepared as described above. The vials were covered with a Teflon pinhole sheet, a silicone pinhole mat, and a steel gas diffusion plate. The reactor insert was placed in a pressure vessel and purged 2x with NH₃ gas. The pressure vessel was charged to 100 psi (689 kPa) with NH₃ gas and then to 680 psi (4688 kPa) with N₂ at ambient temperature. The reactor was placed on a shaker and vortexed at 800 rpm at 160°C. After 3 hours, the reactor was cooled to room temperature, vented, and purged with

nitrogen prior to being unsealed. The samples were diluted with water, mixed, and then centrifuged to separate catalyst particles. Aliquots were removed from the supernatant and diluted further with dilute aqueous trifluoroacetic acid for analysis by HPLC. Results are summarized below in Table 25.

Table 25.

Support	Surface Area (m ² /g)	Mean Pore Diameter (Å)	Pore Volume (cm ³ /g)	Catalyst Amount (mg)	HDO Conversion (%)	HMDA Yield (%)	Pentylamine Yield (%)	Selectivity
Asbury 5346R	80	145	0.28	10	79.8	16.5	0.0	99.8
Asbury 5346R	80	145	0.28	20	99.1	25.1	0.5	97.9
Ensaco 150G	47	136	0.17	10	80.8	16.4	0.1	99.5
Ensaco 150G	47	136	0.17	20	95.5	24.5	0.7	97.4
Ensaco 250G	64	140	0.24	10	78.6	16.3	0.0	100
Ensaco 250G	64	140	0.24	20	94.4	26.0	0.5	98.2
Ensaco 260G	63	104	0.18	10	74.7	15.4	0.0	100
Ensaco 260G	63	104	0.18	20	93.3	24.0	0.5	98.1
HP160	158	208	0.87	10	81.7	11.9	0.0	100
HP160	158	208	0.87	20	98.3	18.3	0.1	99.2
Orion HIBlack 50L	193	157	0.66	10	91.0	18.8	0.2	99.1
Orion HIBlack 50L	193	157	0.66	20	100	21.7	0.7	96.8
Sid Richardson SC159	234	182	0.81	10	90.3	14.7	0.0	100
Sid Richardson SC159	234	182	0.81	20	99.3	17.3	0.4	97.9
Sid Richardson SR155	146	222	0.87	10	93.2	19.5	0.1	99.5
Sid Richardson SR155	146	222	0.87	20	100	19.7	0.4	98.1

Experiment 2

Preparation of Supported Ru/Re Catalysts

[0185] Suitably concentrated aqueous solutions of $\text{Ru}(\text{NO})(\text{NO}_3)_3$ containing varying amounts of HReO_4 were added to 0.15 g of a support and agitated to impregnate the support. The volume of metal solution was matched to equal the pore volume of the support. The samples were dried in an oven at 60°C for 3 hours under a dry air purge. Catalyst amounts from 10-20 mg were weighed in to glass vials of a 96 vial array. The catalysts were reduced under forming gas (5% H_2 and 95% N_2) at 60°C for 3 hours then at 250°C for 3 hours using a $2^\circ\text{C}/\text{min}$ temperature ramp rate. The final catalysts were composed of 4.04 weight percent ruthenium containing various rhenium loadings of 0, 0.4, 0.7, and 1.9 wt.%.

Catalyst Screening Procedure

[0186] A substrate solution consisting of 1.549M 1,6-hexanediol in concentrated aqueous NH_4OH was added to an array of catalysts prepared as described above. The vials were covered with a Teflon pinhole sheet, a silicone pinhole mat, and a steel gas diffusion plate. The reactor insert was placed in a pressure vessel and purged 2x with NH_3 gas. The pressure vessel was charged to 100 psi (689 kPa) with NH_3 gas and then to 680 psi (4688 kPa) with N_2 at ambient temperature. The reactor was placed on a shaker and vortexed at 800 rpm at 160°C . After 3 hours, the reactor was cooled to room temperature, vented, and purged with nitrogen prior to being unsealed. The samples were diluted with water, mixed, and then centrifuged to separate catalyst particles. Aliquots were removed from the supernatant and diluted further with dilute aqueous trifluoroacetic acid for analysis by HPLC. Results are outline below in Table 26.

Table 26. Hexanediol to Hexamethylenediamine using Ru/Re/Carbon HP-160 Catalysts

Entry	Catalyst Amount (mg)	Ru (wt.%)	Re (wt.%)	HDO Conversion (%)	HMDA Yield (%)	Pentylamine Yield (%)	HMDA / Pentylamine
1	20	4.04	0.0	95.8	20.2	1.4	14.5
2	20	4.04	0.4	98.7	23.2	1.3	18.1
3	20	4.04	0.7	99.4	23.4	1.0	23.3
4	20	4.04	1.9	97.4	20.3	0.4	51.8
5	10	4.04	0.0	79.5	14.4	0.6	26.1
6	10	4.04	0.4	90.4	20.1	0.7	27.9
7	10	4.04	0.7	92.6	20.7	0.6	35.8
8	10	4.04	1.9	78.3	11.5	0.0	

Experiment 3

Preparation of Ni/Ru Catalysts Supported on Ensaco 250G

[0187] Suitably concentrated aqueous solutions containing $\text{Ni}(\text{NO}_3)_2$ and/or $\text{Ru}(\text{NO})(\text{NO}_3)_3$ were added by incipient wetness impregnation to approximately 0.4 g of a carbon black support and agitated to

impregnate the support. The volume of metal solution was matched to equal the pore volume of the support. Each catalyst was thermally treated under N₂ in a tube furnace at 60°C for 12 hours then at 300°C for 3 hours using a 5°C/min temperature ramp rate.

[0188] Catalyst amounts from 15-25 mg were weighed in to glass vials of a 96 vial array. The catalysts were reduced under forming gas (5% H₂ and 95% N₂) at 450°C for 3 hours using a 2°C/min temperature ramp rate. Catalysts were passivated with 1% O₂ in N₂ at room temperature before removing from the tube furnace.

Catalyst Screening Procedure A

[0189] A substrate solution consisting of 0.7M 1,6-hexanediol in concentrated aqueous NH₄OH was added to an array of catalysts prepared as described above. The vials were covered with a Teflon pinhole sheet, a silicone pinhole mat, and a steel gas diffusion plate. The reactor insert was placed in a pressure vessel and purged 2x with NH₃ gas. The pressure vessel was charged to 100 psi (689 kPa) with NH₃ gas and then to 680 psi (4688 kPa) with N₂ at ambient temperature. The reactor was placed on a shaker and vortexed at 800 rpm at 160°C. After 3 hours, the reactor was cooled to room temperature, vented, and purged with nitrogen prior to being unsealed. The samples were diluted with water, mixed, and then centrifuged to separate catalyst particles. Aliquots were removed from the supernatant and diluted further with dilute aqueous trifluoroacetic acid for analysis by HPLC. Results are summarized below in Table 27.

Table 27.

Entry	Catalyst Amount (mg)	Ni (wt.%)	Ru (wt.%)	HDO Conversion (%)	HMDA Yield (%)	Pentylamine Yield (%)	Selectivity
1	15	12	0	35.8	1.9	0.0	98.4
2	15	0	0.56	66.1	9.6	0.3	97.4
3	15	4	0.56	74.3	17.3	0.3	98.5
4	15	8	0.56	77.3	22.7	0.3	98.7
5	15	12	0.56	85.9	25.3	0.4	98.6
6	25	12	0	54.0	7.7	0.1	99.1
7	25	0	0.56	82.7	18.9	0.5	97.2
8	25	4	0.56	89.1	29.0	0.8	97.5
9	25	8	0.56	92.8	31.4	0.7	97.7
10	25	12	0.56	94.7	30.1	0.7	97.6

Catalyst Screening Procedure B

[0190] Passivated catalysts were reactivated in water under H₂ at 180°C for 3 hours. Most of the water was removed from each catalyst leaving behind enough to act as a protective layer. The catalysts were then screened as described above in Procedure A. Results are summarized below in Table 28.

Table 28.

Entry	Catalyst Amount (mg)	Ni (wt.%)	Ru (wt.%)	HDO Conversion (%)	HMDA Yield (%)	Pentylamine Yield (%)	Selectivity
1	15	12	0	49.2	7.2	0.00	100.0
2	15	0	0.56	67.0	11.2	0.06	99.5
3	15	4	0.56	73.7	18.2	0.07	99.7
4	15	8	0.56	86.8	25.5	0.38	98.5
5	15	12	0.56	86.0	23.8	0.33	98.7
6	25	12	0	73.3	16.0	0.00	100.0
7	25	0	0.56	83.8	18.3	0.84	95.6
8	25	4	0.56	93.5	28.1	1.22	95.9
9	25	8	0.56	94.0	27.8	1.07	96.3
10	25	12	0.56	96.1	27.3	1.15	96.0

Fixed Bed Experiments

Preparation of 2 wt.% Ru on Carbon Ensaco 250G

[0191] A carbon extrudate, prepared from carbon black Ensaco 250G and a carbohydrate binder, was crushed and sized to 150-300um. A suitably concentrated aqueous solution of $\text{Ru}(\text{NO})(\text{NO}_3)_3$ was added to 4.77 g of the crushed extrudate and agitated to impregnate the support. The volume of metal solution was matched to equal the pore volume of the support. The samples were dried in an oven at 60°C for 12 hours under a dry air purge. The catalyst was reduced under forming gas (5% H_2 and 95% N_2) at 250 °C for 3 hours using a 2°C/min temperature ramp rate. The catalyst was washed with water and again sized to 106-300um to remove any fines that may have been generated during the metal impregnation step.

Preparation of 10.5 wt.% Ni and 0.45 wt.% Ru on Carbon Ensaco 250G

[0192] A carbon extrudate, prepared from carbon black Ensaco 250G and a carbohydrate binder, was crushed and sized to 106-300um. A suitably concentrated aqueous solution containing $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and $\text{Ru}(\text{NO})(\text{NO}_3)_3$ was added to 10 g of the crushed extrudate and agitated to impregnate the support. The volume of metal solution was matched to equal the pore volume of the support. The catalyst was dried in an oven at 60°C for 12 hours under a dry air purge then thermally treated under N_2 at 300°C for 3 hours. The catalyst was reduced under forming gas (5% H_2 and 95% N_2) at 450°C for 3 hours using a 2°C/min temperature ramp rate. After cooling to room temperature the catalyst was passivated with 1% O_2 in N_2 at room temperature before removing from the tube furnace. The catalyst was washed with water and again sized to 106-300um to remove any fines that may have been generated during the metal impregnation step.

2 wt.% Ru on Carbon catalyst

[0193] The reaction was performed in a 0.25 inch OD by 570 mm long 316 stainless steel tube with a 2 um 316 stainless steel frit at the bottom of the catalyst bed. The reactor was vibration packed with 1g of SiC beads (90-120um) followed by 3g of a 2% by weight ruthenium on carbon Ensaco 250G catalyst (100-300 um) and finally 2.5g of SiC beads at the top. A ¼ inch layer of glass wool was used between each layer. The packed reactor tube was vertically mounted in an aluminum block heater equipped with PID controller. An HPLC pump was used to deliver liquid feed to the top of the reactor and a back pressure regulator was used to control reactor pressure. The reaction was run at 160°C. Product effluent was collected periodically for analysis by HPLC. No decline in catalyst activity was observed after 1650h.

[0194] Three different feed compositions were investigated at 160°C with a reactor pressure ranging from 800-1000 psi (5516 6895 kPa). In all cases N-methyl-2-pyrrolidone (NMP) was used as an internal standard.

Feed 1: 0.7M 1,6-hexanediol and 0.14M NMP in concentrated NH4OH.

Feed 2: 0.7M 1,6-hexanediol, 0.14M hexamethyleneimine, and 0.14M NMP in concentrated NH4OH.

Feed 3: 1.54M 1,6-hexanediol, 0.308M hexamethyleneimine, and 0.308M NMP in concentrated NH4OH.

[0195] The results are summarized below in Table 29.
10.5 wt.% Ni /0.45 wt.% Ru on Carbon catalyst

[0196] The reaction was performed as described above for the Ru only catalyst. A total of 3g of Ni/Ru catalyst was loaded into the reactor and reactivated at 180°C under H₂ before introduction of the feed solution. No decline in catalyst activity was observed after 650h. Results are summarized below in Table 29.

Table 29.

Catalyst	Source	Feed / Rate (mL/mm)	Reactor Pressure (psi)	HDO Conversion (%)	HMDA Yield (%)	HMI Yield (%)	Pentylamine Yield (%)	Selectivity
2wt.% Ru	Feed 1	0.2	800	85	23	11	1.5	93.9
2wt.% Ru	Feed 2	0.2	800	83	29	18	1.6	94.8
2wt.% Ru	Feed 3	0.15	1000	90	36	21	2.6	93.3
10.5wt.% Ni/0.45%Ru	Feed 2	0.2	800	83	26	17	0.5	98.1

[0197] When introducing elements of the present invention or the preferred embodiments(s) thereof, the articles "a", "an", "the" and "said" are intended to mean that there are one or more of the elements. The terms "comprising", "including" and "having" are intended to be inclusive and mean that there may be additional elements other than the listed elements.

[0198] In view of the above, it will be seen that the several objects of the invention are achieved, and other advantageous results attained.

REFERENCES CITED IN THE DESCRIPTION

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Patentkrav

1. Fremgangsmåde til fremstilling af en katalysatorsammensætning, hvilken fremgangsmåde omfatter:

- 5 at blande vand, carbon black, og et vandopløseligt organisk bindemiddel, i fraværet af et organisk opløsningsmiddel, at danne en carbon black blanding, hvor bindemidlet omfatter: (i) et saccharid valgt fra gruppen bestående af et monosaccharid, et disaccharid, et oligosaccharid og en hvilken som helst kombination deraf og (ii) et polymert kulhydrat, hvor vandindholdet af carbon black blandingen er fra 5 vægtprocent til 70
- 10 vægtprocent, carbon black indholdet af carbon black blandingen er mindst 35 vægtprocent på tørvægtsbasis, og bindemiddelindholdet af carbon black blandingen er mindst 10 vægtprocent på tørvægtsbasis, og hvor et bindemiddel anses for vandopløseligt, hvis opløseligheden ved 50°C er mindst ca. 1 vægtprocent;
- 15 at danne carbon black blandingen for at fremstille et formet carbon black kompositmateriale;
- at opvarme det formede carbon black kompositmateriale ved en temperatur fra 300°C til 800°C for at karbonisere bindemidlet til en vanduopløselig tilstand og for at fremstille et formet porøst carbonprodukt;
- 20 og
- at aflejre en katalytisk aktiv komponent eller prækursor deraf omfattende et ædelmetal på det formede porøse carbonprodukt for at fremstille katalysatorsammensætningen.

25 **2.** Fremgangsmåden ifølge krav 1, hvor vægtforholdet mellem bindemidlet og carbon black i carbon black blandingen er fra 1:4 til 1:1.

3. Fremgangsmåden ifølge krav 1 eller 2, hvor ædelmetallet er valgt fra gruppen bestående af ruthenium, rhodium, palladium, platin og kombinationer deraf.

30

4. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 3, hvor carbon black indholdet af carbon black blandingen er fra 35 vægtprocent til 80

vægtprocent, og/eller koncentrationen af bindemidlet i carbon black blandingen er fra 10 vægtprocent til 50 vægtprocent på tørvægtbasis.

5. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 4, hvor
5 bindemidlet omfatter et monosaccharid valgt fra gruppen bestående af glucose, fructose, hydrat deraf, sirup deraf og kombinationer deraf.

6. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 5, hvor det
10 polymere kulhydrat omfatter en celluloseforbindelse valgt fra gruppen bestående af methylcellulose, ethylcellulose, ethylmethylcellulose, hydroxyethylcellulose, hydroxypropylcellulose, methylhydroxyethylcellulose, ethylhydroxyethylcellulose, hydroxypropylmethylcellulose, carboxymethylcellulose og blandinger deraf.

7. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 6, hvor
15 bindemidlet omfatter et saccharid valgt fra gruppen bestående af glucose, fructose eller hydrat deraf og et polymert kulhydrat valgt fra gruppen bestående af hydroxyethylcellulose, methylcellulose og stivelse.

8. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 7, hvor carbon
20 black materialet har et specifikt BET-overfladeareal fra 20 m²/g til 500 m²/g.

9. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 8, hvor carbon
black materialet har en gennemsnitlig porediameter fra 5 nm til 100 nm som målt med BJH-metoden.

25

10. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 9, hvor carbon
black materialet har en specifik porevolumen fra 0,1 cm³/g til 1 cm³/g som målt med BJH-metoden.

30 **11.** Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 10, hvor det formede porøse carbonkompositmateriale dannes ved ekstrudering af carbon black blandingen.

12. Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 11, hvor carbon black blandingen dannes under et tryk fra 100 kPa (1 bar) til 10.000 kPa (100 bar).

5 **13.** Fremgangsmåden ifølge et hvilket som helst af kravene 1 til 12, hvor ædelmetallet omfatter platin.

DRAWINGS

FIG. 1

Cross Sectional Analysis of a 1.5 mm Carbon Black Extrudate Catalyst

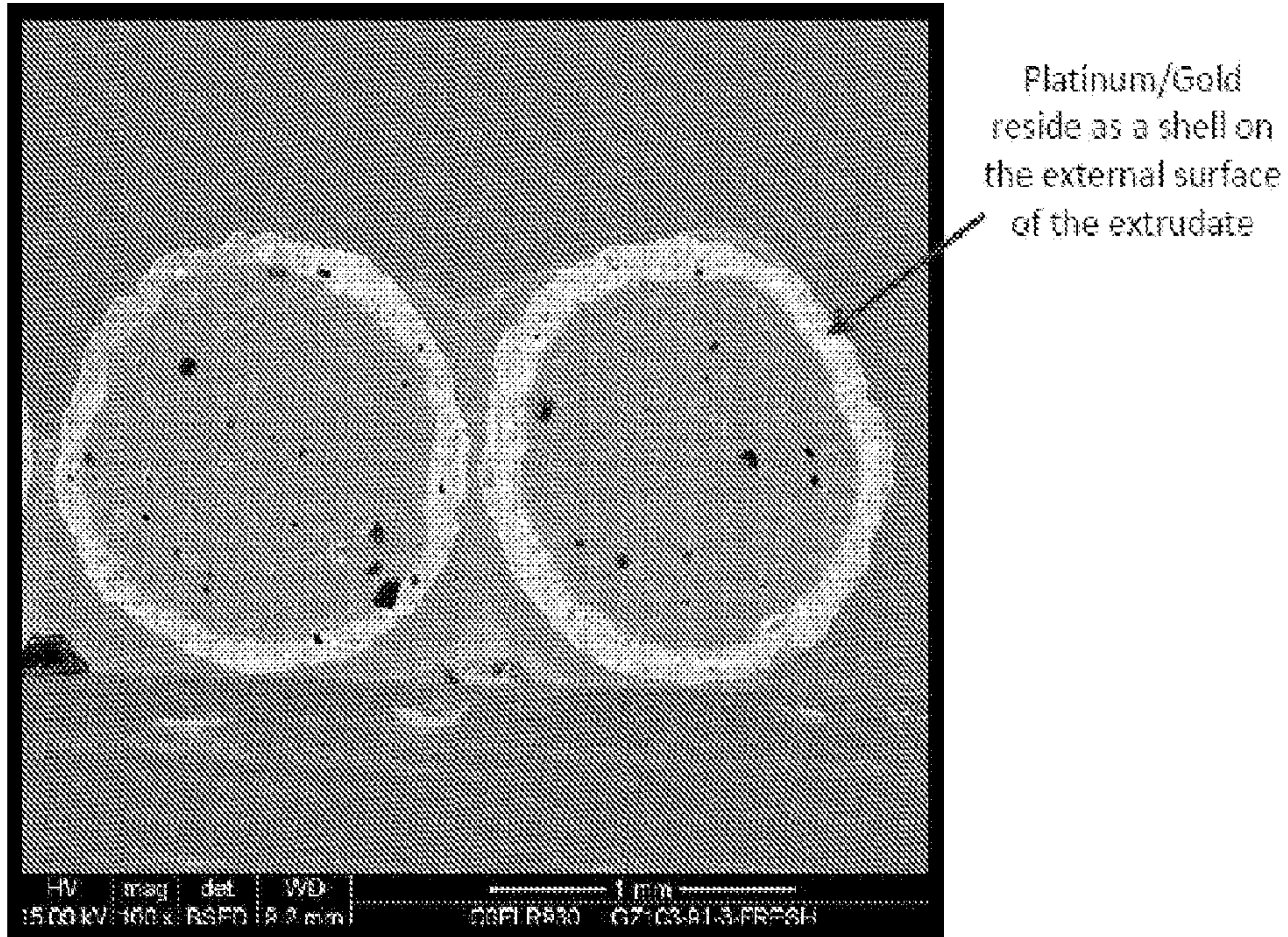


FIG. 2

Cross Sectional Analysis of a 1.5 mm Carbon Black Extrudate Catalyst

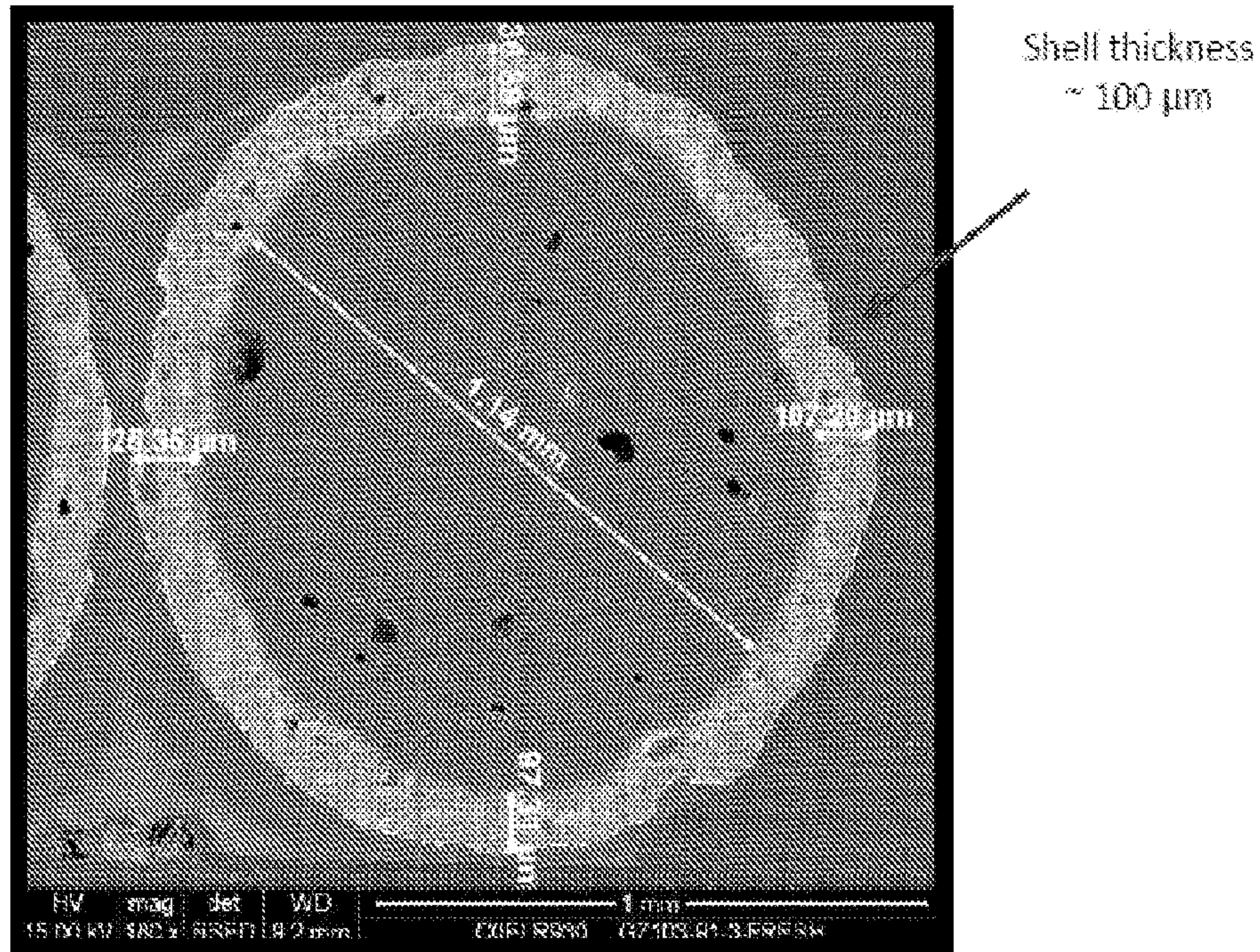


FIG. 3

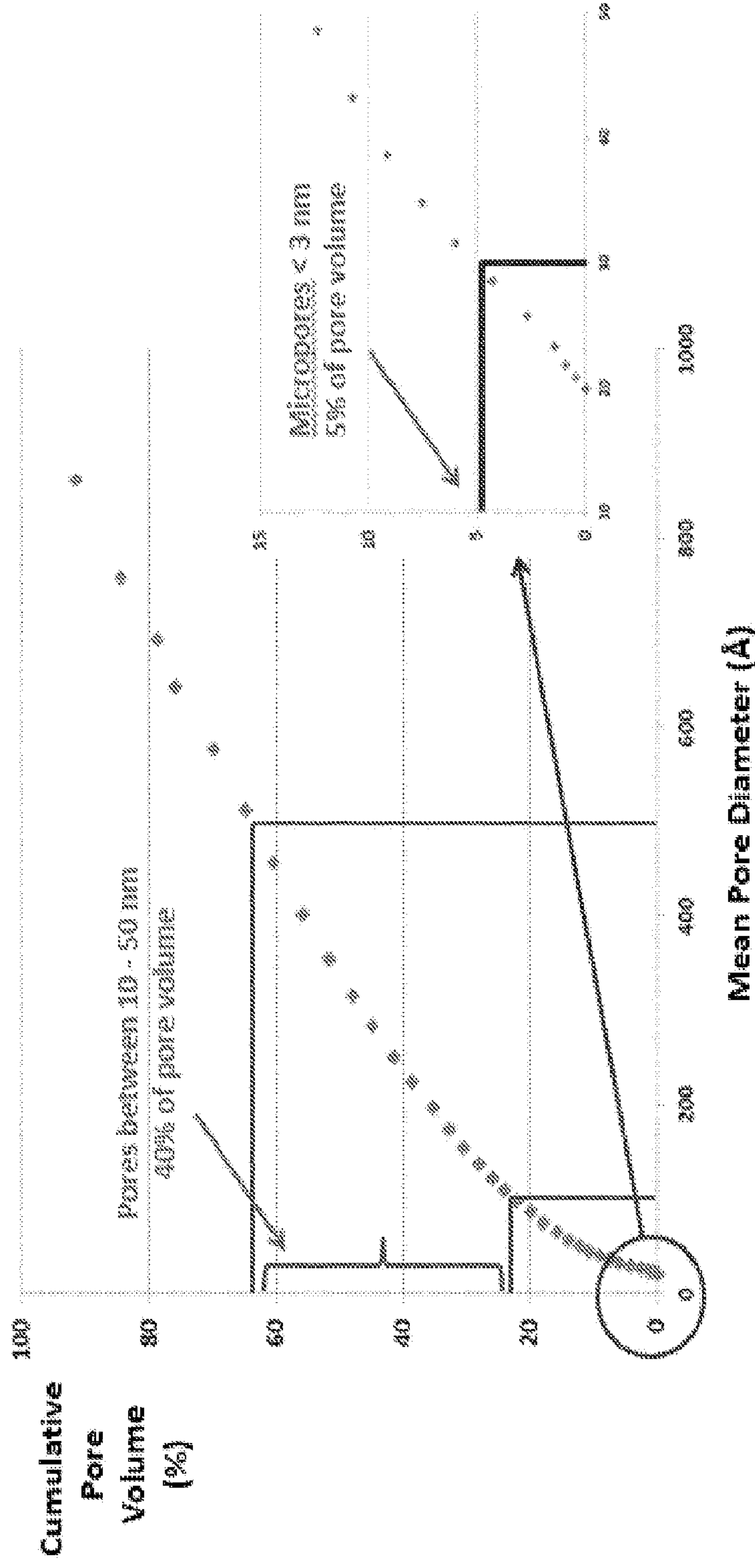


FIG. 4

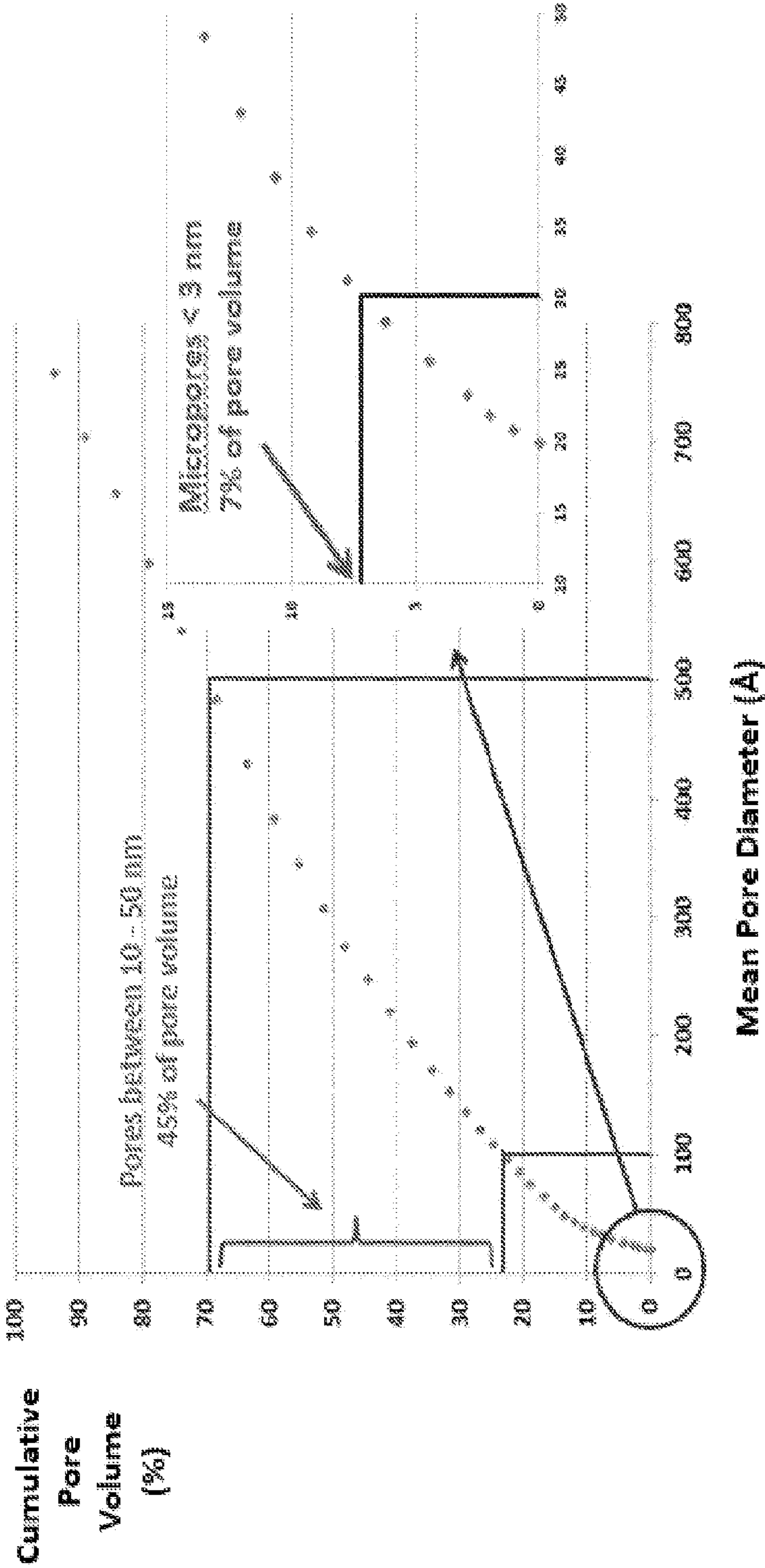


FIG. 5

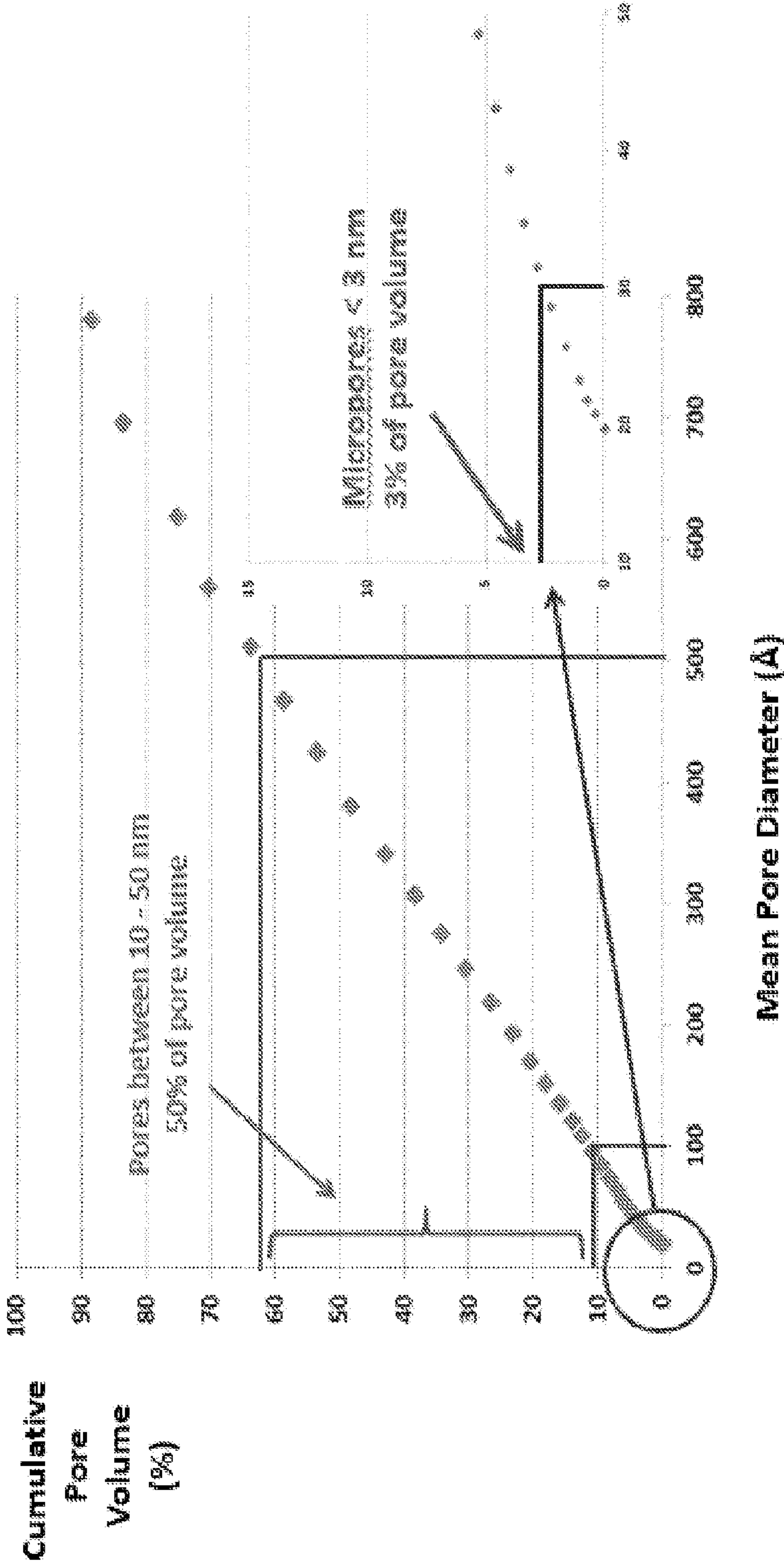


FIG. 6

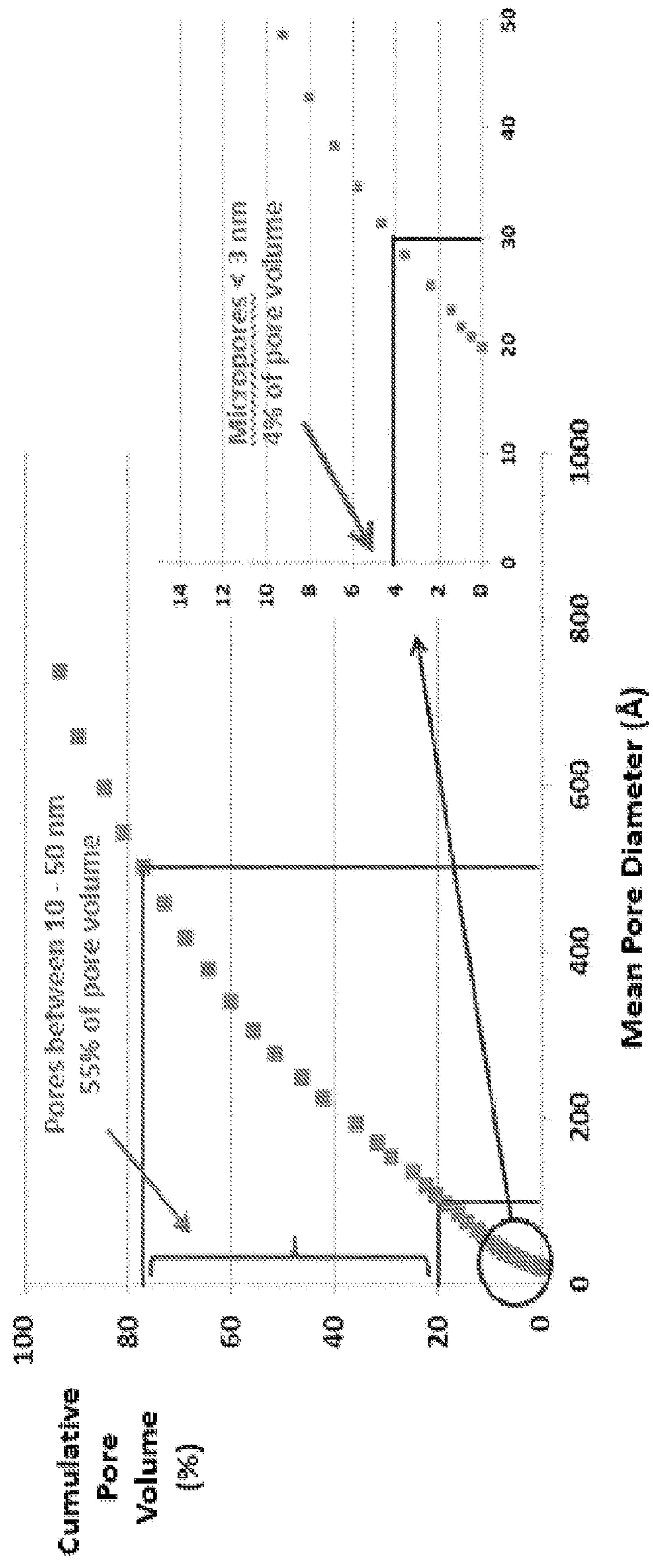


FIG. 7

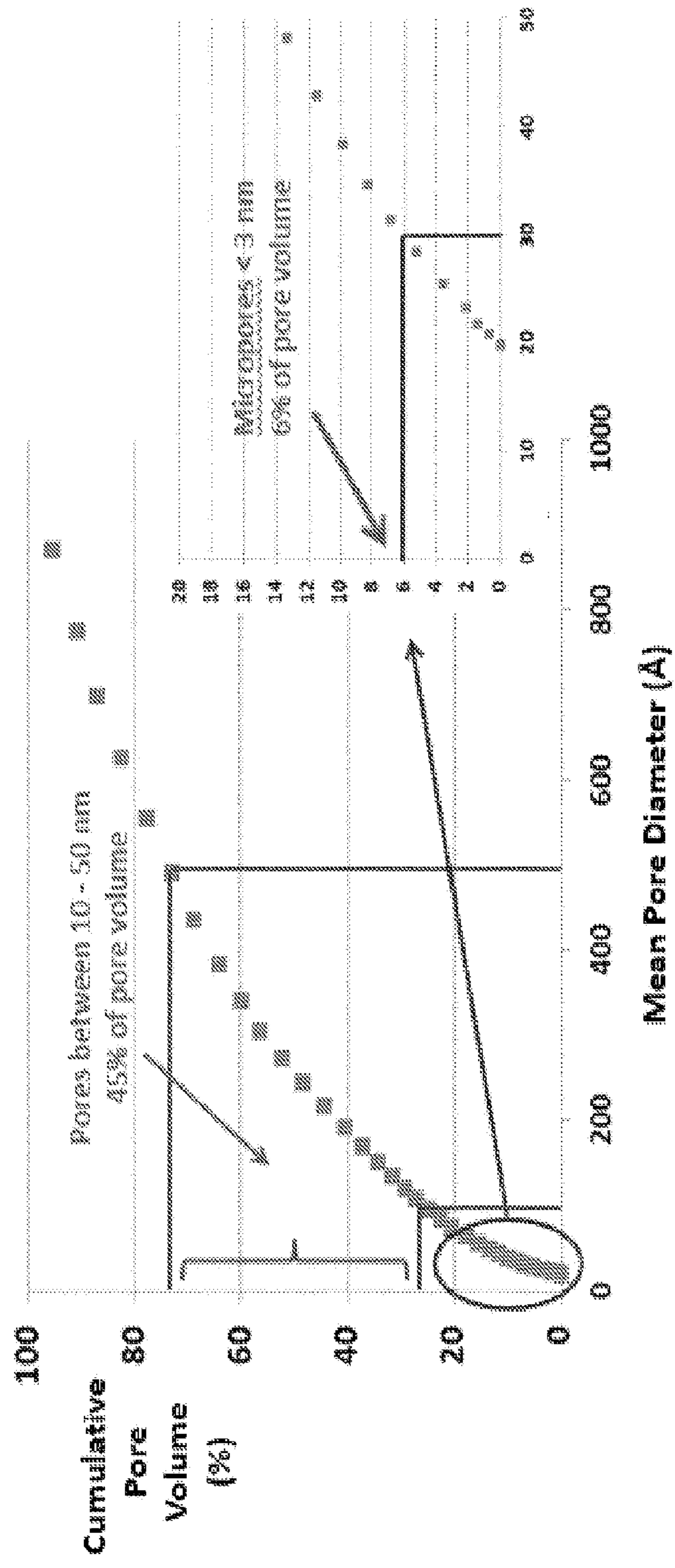


FIG. 8

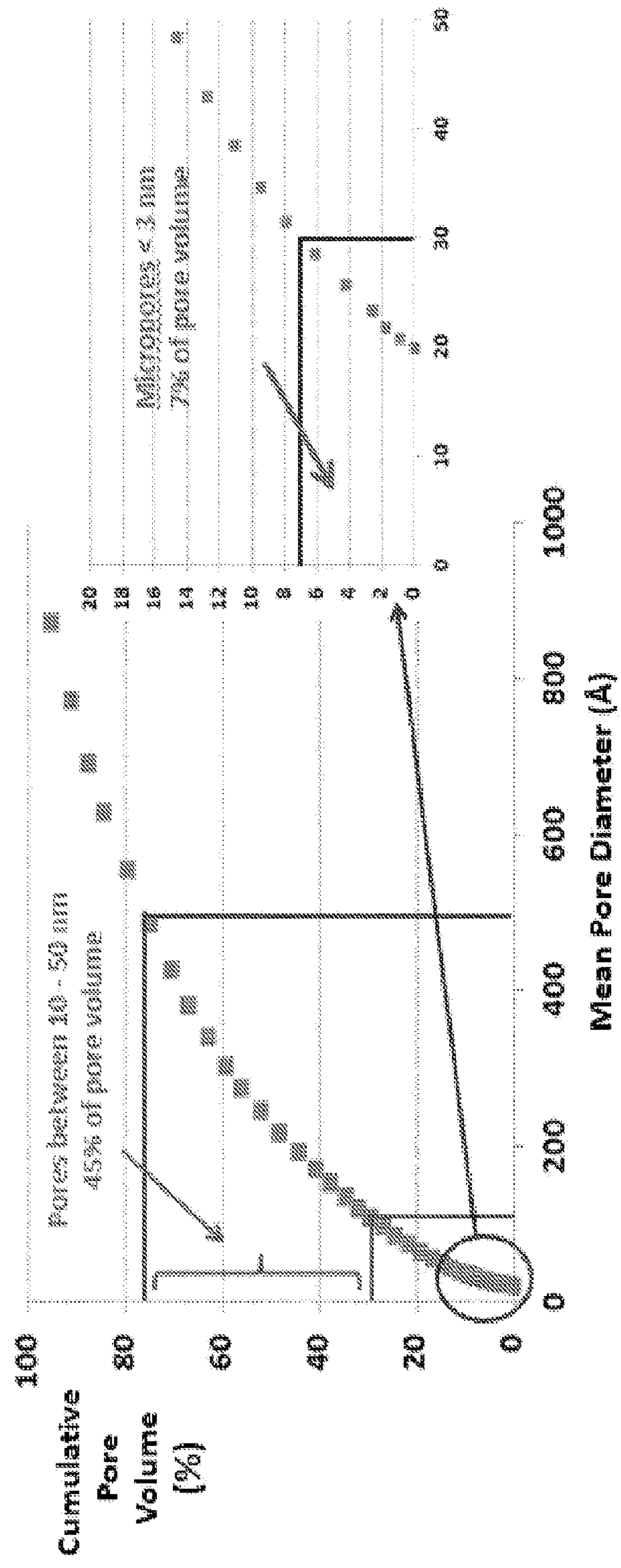


FIG. 9

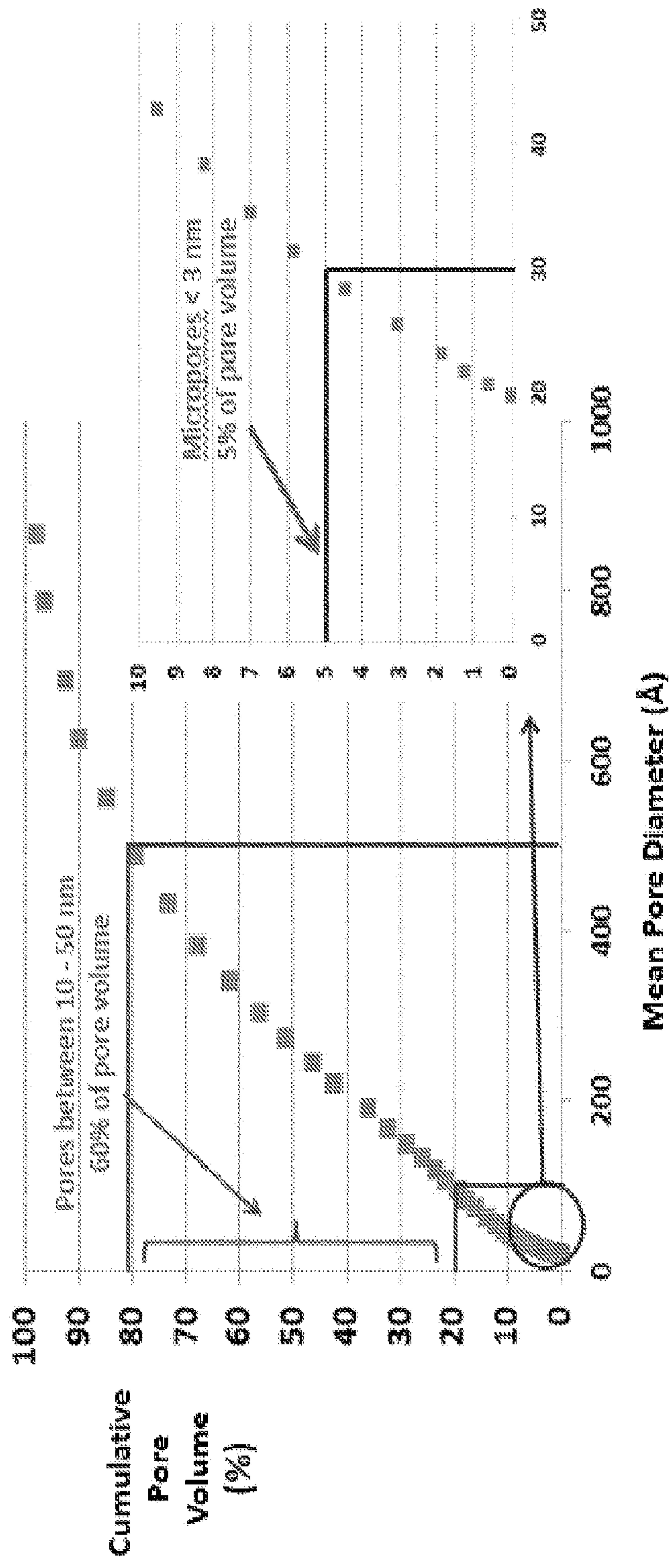


FIG. 10

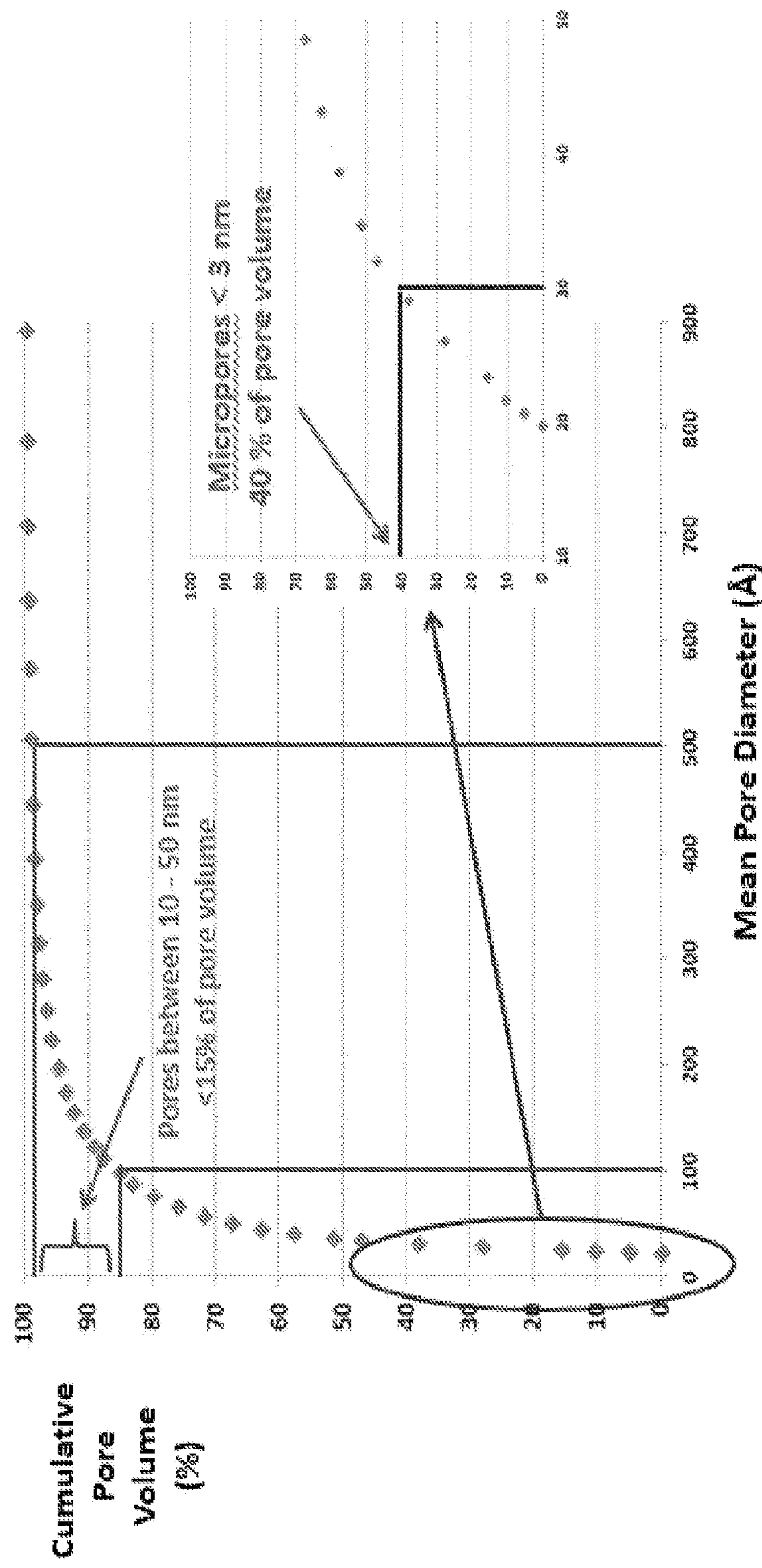


FIG. 11

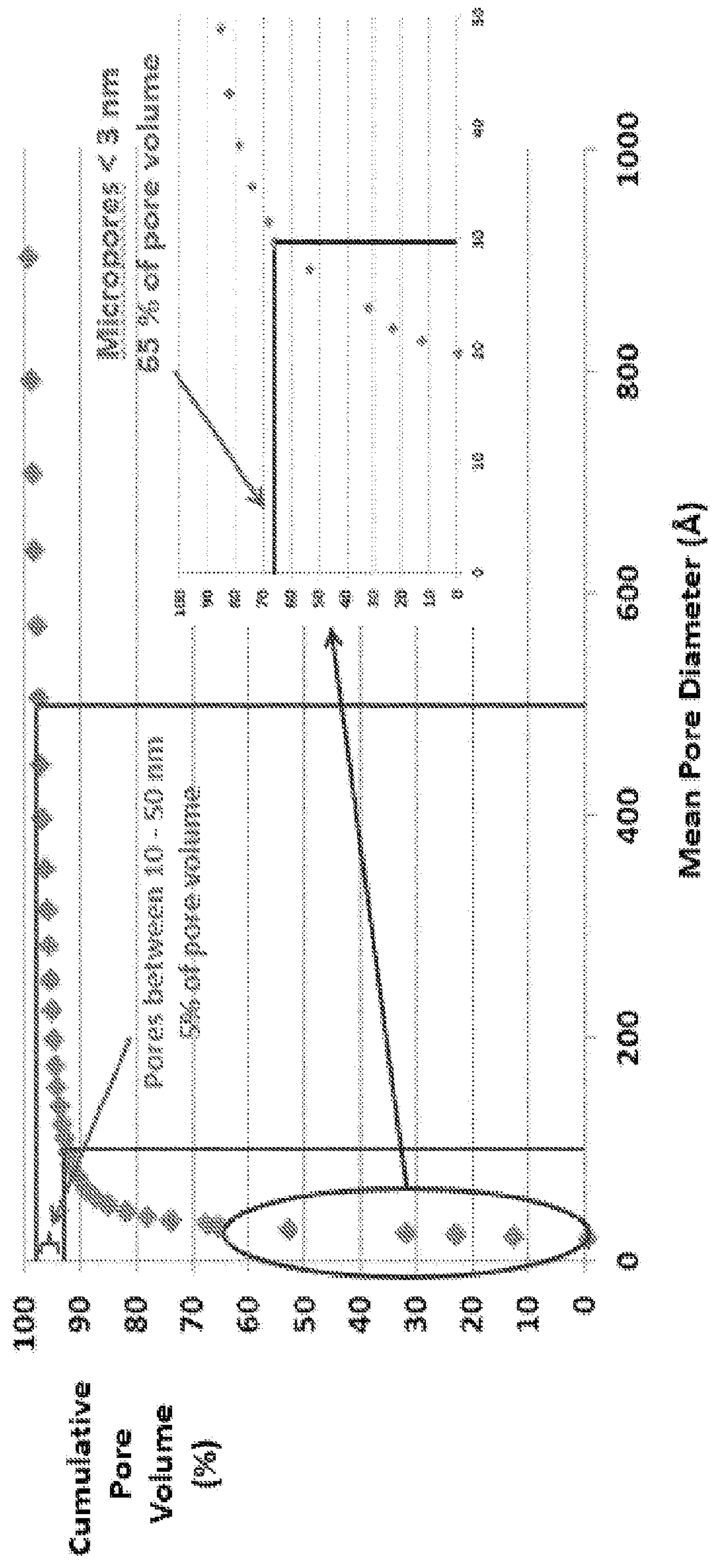


FIG. 12

