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(54) Titre : COMPOSITION DE REVETEMENT DURCISSABLE OXYDATIVEMENT
(54) Title: OXIDATIVELY CURABLE COATING COMPOSITION

(57) Abrégé/Abstract:

The present invention relates to an oxidatively curable coating formulation comprising an oxidatively curable alkyd-based resin and a diazacycloalkane-based chelant, which chelant may optionally be complexed with a suitable transition metal ion. The formulations may be paints or other oxidatively curable coating compositions. The invention also provides methods for making such formulations and compositions resultant from the curing of such formulations.

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(54) Title: OXIDATIVELY CURABLE COATING COMPOSITION

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OXIDATIVELY CURABLE COATING COMPOSITION

FIELD

The present invention relates to an oxidatively curable coating formulation comprising an oxidatively curable alkyd-based resin and a diazacycloalkane-based chelant, which chelant may optionally be complexed with a suitable transition metal ion. The formulations may be paints or other oxidatively curable coating compositions. The invention also provides methods for making such formulations and compositions resultant from the curing of such formulations.

10

BACKGROUND

Alkyd resins are a well understood and dominant binder in many oxidatively curable paints and other solvent-based coatings. Alkyd emulsion paints, in which the continuous phase is aqueous, are also widely available commercially. Alkyd resins are produced by the reaction of polyols with carboxylic acids or anhydrides. To make them susceptible to what is commonly referred to as a drying process, some alkyd resins are reacted with unsaturated triglycerides or other source of unsaturation. Plant and vegetable oils, such as linseed oil, are frequently used as the source of triglycerides. In these drying processes, unsaturated groups, in particular carbon-carbon double bonds, can react with oxygen from the air, causing the oils to crosslink, forming a three-dimensional network, and harden. This oxidative curing process, although not drying, gives the appearance of drying and is often and herein referred to as such. The length of time required for drying depends on a variety of factors, including the constituents of the alkyd resin formulation and the amount and nature of the liquid continuous phase (e.g. solvent) in which the alkyd resin is formulated.

Film formation results from the autoxidation and polymerisation chemistries that occur during the drying of alkyd-based resins. It will proceed in the absence of catalysis. However, it is customary to include in formulations of curable resins small, i.e. catalytic, quantities of optionally organic metal salts, often referred to as metal driers, which catalyse the polymerisation of unsaturated material so as to form the three-dimensional network.

Driers used for solvent-based coatings typically include alkyl carboxylates, typically C₆-C₁₈ carboxylates, of metals such as cobalt, manganese, lead, zirconium, zinc, vanadium, strontium, calcium and iron. Such metal carboxylates are often referred to as metal soaps. Redox-active metals, such as manganese, iron, cobalt, vanadium and copper enhance radical formation, and thus the oxidative curing process, whilst so-called secondary driers (sometimes referred to as auxiliary driers),

such as complexes based on strontium, zirconium and calcium, enhance the action of the redox-active metals. Often these soaps are based on medium-chain alkyl carboxylates such as 2-ethyl-hexanoate. The lipophilic units in such soaps enhance the solubility of the drier in solvent-based paints and other oxidatively curable coating compositions.

As well as metal soaps, a variety of metal driers that are redox metal complexes containing organic ligands can be used as driers, for example manganese complexes comprising 2,2'-bipyridine or 1,10-phenanthroline ligands.

The formation of a skin or lumpy matter is a problem observed in many oil-based (i.e. organic solvent-based) formulations, and in particular in organic solvent-based alkyd resins, as a consequence of oxidation during storage or transportation. Oxidative polymerisation reactions can thus lead to the skin formation before application, as well as the intended drying after application. As alluded to above, these polymerisation reactions can be triggered by radicals generated by the action of metal-based driers, for example cobalt-, manganese- or iron-containing driers. In other words, the cause of the skin formation is often associated with the presence of metal driers.

Skin formation during manufacture and storage of air-drying paints and other coatings, in particular of alkyd-based resins, is clearly undesirable. Skin formation can lead to material losses and usage problems, such as surface irregularity after application owing to skin particles remaining in the paint.

Addition of compounds that quench the radicals formed during the storage or transportation processes reduce the skin-forming tendencies of such formulations. Many antiskinning agents are therefore antioxidants. However, addition of such antiskinning antioxidants can also slow the drying desired after application, by reducing the activity of the metal driers.

Oximes, and in particular methylethylketoxime (MEKO), are known to reduce skin formation considerably, particularly with cobalt-based driers. It is understood that the oxime binds to the metal ion during storage of the resin, thereby preventing the metal drier from reacting with oxygen and the substrate for radical formation that otherwise leads to polymerisation and skin formation. Upon application of the paint or other coating as a thin layer on a surface, the MEKO can evaporate. In this way, skinning can be prevented or ameliorated, but the cobalt soap can function, after application, as a polymerisation catalyst (see J H Bieleman in *Additives in Plastics and Paints, Chimia*, **56**, 184 (2002)).

Antiskinning agents, or ways to address the problem of skinning, other than those involving the use of oximes such as MEKO, have been described. For example,

WO 00/11090 A1 (Akzo Nobel N.V.) describes the use of 1,3-diketones, pyrazoles and imidazoles to reduce the skinning properties; WO 2007/024592 A1 (Arkema Inc.) describes the use of isoascorbate as an antiskinning agent and a co-promoter of metal-based driers; and WO 2008/127739 A1 (Rockwood Pigments NA, Inc.) describes the use of hydroxylamine as an antiskinning agent. Whilst such additives reduce the tendency towards skinning, they can lead to decreased performance of the metal drier if their degree of incorporation is too great and they do not evaporate sufficiently during the coating (e.g. paint) application.

Whilst cobalt driers have been employed for many years as paint driers, there is a desire to develop alternatives, not least since cobalt soaps may need to be registered as carcinogenic materials. Iron- and manganese-based paint driers in particular have received considerable attention in recent years in the academic and patent literature as alternatives to cobalt-based driers. For some recent scientific publications addressing this topic in detail see publications by J H Bieleman (in *Additives in Plastics and Paints, Chimia, infra*); J H Bieleman (*Marcomol. Symp.*, **187**, 811 (2002)); and R E van Gorkum and E Bouwman (*Coord. Chem. Rev.*, **249**, 1709 (2005)).

WO 03/093384 A1 (Ato B.V.) describes the use of reducing biomolecules in combination with transition-metal salts or complexes based on pyrazoles, aliphatic and aromatic amines, 2,2'-bipyridine, 1,10'-phenanthroline and 1,4,7-trimethyl-1,4,7-triazacyclononane (Me₃TACN).

WO 03/029371 A1 (Akzo Nobel N.V.) describes the use of complexes comprising Schiff base compounds to enhance the drying of coatings, in which complexes at least one solubilising group is covalently bound to the organic ligand.

EP 1382648 A1 (Universiteit Leiden) describes the use of manganese complexes with acetylacetonate and bidentate nitrogen donor ligands in paint drying.

WO 2008/003652 A1 (Unilever PLC *et al.*) describes the use of tetradentate, pentadentate or hexadentate nitrogen ligands bound to manganese and iron as siccative for curing alkyd-based resins.

Oyman *et al.* describe the oxidative drying of alkyd paints by [Mn₂(μ-O)₃(Me₃tacn)₂](PF₆)₂ (Z O Oyman *et al.*, *Surface Coating International Part B – Coatings Transaction*, **88**, 269 (2005)). WO 2011/098583 A1, WO 2011/098584 A1 and WO 2011/098587 A1 (each DSM IP Assets B.V.) describe the use of a variety of dinuclear manganese complexes with Me₃TACN as ligand for paint drying.

WO 2012/092034 A2 (Dura Chemicals, Inc.) describes the use of a transition metal and a porphyrin based ligand as a siccative for resin compositions.

The use of mixtures of metal salts and ligands to enhance drying of paint formulations is known. For example, W H Canty, G K Wheeler and R R Myers (*Ind.*

Eng. Chem., **52**, 67 (1960)) describe the drying capability of a mixture of 1,10-phenanthroline (phen) and Mn soap, which is similar to that of prepared Mn-phen complexes. Mixtures of 2,2'-bipyridine (bpy) and manganese soaps show a better drying performance than manganese soaps without bpy (see P K Weissenborn and A Motiejauskaite, *Prog. Org. Coat.*, **40**, 253 (2000)). Also, R van Gorkum *et al.* (*Inorg. Chem.*, **43**, 2456 (2004)), describe that the addition of bpy to Mn(acetylacetonate)₃ gives an acceleration in the drying performance, and attribute this to the formation of manganese-bipyridine complexes. The use of manganese complexes with acetylacetonate and bidentate nitrogen donor ligands in paint drying has also been described in EP 1382648 A1 (Universiteit Leiden).

In WO 2012/093250 A1 (OMG Additives Limited) it is described that, by contacting an aqueous solution of transition metal ions and polydentate ligands with alkyd-based formulations, the resultant formulation shows reduced skinning tendency as compared with the introduction of metal ions and polydentate ligands in nonaqueous media.

It may be inferred from the recent literature, including patent literature, published in the field of oxidatively curable coating formulations, for example from WO 2008/003652 A1, WO 2011/098583 A1, WO 2011/098584 A1, WO 2011/098587 A1 and WO 2012/092034 A2, that advantageous curing rates of oxidatively curable resins, for example alkyd-based resins, results from the use of metal driers comprising ligands that give rise to relatively stable transition metal-ligand complexes. In general, when using polydentate ligands, i.e. ligands that bind a metal ion through more than one donor site, improved stability of the resultant metal complexes in different redox states can be observed as compared to the corresponding complexes where monodentate ligands are used.

Triazacyclononane- (TACN-) based ligands in particular can be used to provide stable complexes with a variety of transition metal ions. The use of macrocyclic ligands such as TACN-based ligands is believed to afford complexes with greater thermodynamic stabilities as compared to the use of linear triamines, for example. Further, the 9-membered ring size of TACN-based ligands is believed to give rise to further enhanced thermodynamic stabilities, as compared to, for example, use of ligands with 10-, 11- or 12-membered ring systems having 3 tertiary nitrogen donor atoms. More extensive information on these considerations can be found in P Chaudhuri and K Wieghardt (*Prog. Inorg. Chem.*, **35**, 329 (1987)).

WO 01/85717 A1 (Unilever plc *et al.*) exemplifies the use of diazacycloalkane-based ligands with manganese, iron and cobalt salts as stain-bleaching catalysts for use in bleaching. Whilst significant curry stain bleaching activity was noted, the effect

on tea stains was quite low and much lower than that found with manganese complexes of 1,4,7-trimethyl-1,4,7-triazacyclononane. One of the reasons for this difference may be attributable to the greater stability of TACN-based complexes over the diazacycloalkane-containing complexes described in this publication, where only two of the coordinating nitrogen donor atoms are part of the diazacycloalkane, with additional nitrogen donor atoms not being part of the TACN ring.

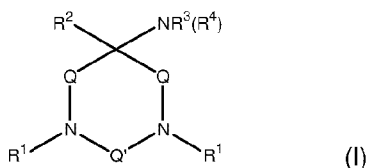
A Neves *et al.* (*Inorg. Chem.*, **44**, 7690 (2005)) reported a comparison of the binding constants found in complexes of nickel comprising diazacycloalkane-based ligands and a TACN analogue, which revealed a difference in 5 log K values.

There remains a need in the art of oxidatively curable formulations for the provision of further curable formulations, which need not comprise cobalt-based driers, but which nevertheless exhibit acceptable rates of curing. Also, there remains a need in the field of oxidatively curable alkyd-based formulations to be able to provide a formulation which, on the one hand, ameliorates the problem of skinning upon storage of such formulations that comprise metal-based driers, and on the other hand requires less modification by the manufacturers of oxidatively curable coating compositions suitable for application than existing oxidatively curable alkyd-based formulations that are essentially absent metal-based driers. The present invention is intended to address these needs.

SUMMARY

We have surprisingly found that transition metal complexes, for example, manganese and iron complexes, comprising diazacycloalkane ligands, catalyse faster curing of oxidatively curable coating formulations than would have been expected given their mixed performance in catalysing bleaching and poorer stabilities than other polydentate ligands.

Viewed from a first aspect, therefore, the invention provides a formulation comprising an oxidatively curable alkyd-based resin and a chelant of the formula (I):



wherein:

each of $-R^1$, $-R^2$, $-R^3$ and $-R^4$ independently represents $-H$, $-C_{1-24}$ alkyl, $-C_{6-10}$ aryl or a group comprising a heteroatom capable of coordinating to a metal ion;

Q represents methylene or ethylene, in which one or more hydrogen atoms may be optionally independently replaced with a C_{1-24} alkyl or a C_{6-10} aryl group; and

Q' represents ethylene or n-propylene, in which one or more hydrogen atoms may be optionally independently replaced with a C₁₋₂₄ alkyl or a C₆₋₁₀ aryl group.

The chelant in the formulation may or may not be part of a complex comprising a suitable transition metal ion.

5 Viewed from a second aspect, the invention provides a method of preparing a formulation according to the first aspect of the invention, the method comprising contacting a composition comprising an alkyd-based resin with a composition comprising a chelant of formula (I). The chelant in the composition comprising it may or may not be part of a complex comprising a suitable transition metal ion.

10 Viewed from a third aspect, the invention provides a composition resultant from curing of a formulation of the first aspect of the invention, or of a formulation obtainable according to the second aspect of the invention.

Viewed from the fourth aspect, the invention provides a chelant of formula (I), which is 6-{N,N-bis(pyridin-2-ylmethyl)amino}-1,4,6-trimethyl-1,4-diazacycloheptane or
15 6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane.

Viewed from a fifth aspect, the invention provides a kit comprising a formulation according to the first aspect of the invention, or obtainable according to the second aspect of the invention, which formulation comprises less than 0.001% by weight of
20 ions of each of at least manganese, iron, cobalt, vanadium and copper and, separately, a composition comprising transition metal ions selected from the group consisting of manganese, iron, cobalt, vanadium and copper ions.

Viewed from a sixth aspect, the invention provides a method comprising applying to a substrate a formulation according to the first aspect of the invention, or
25 obtainable according to the second aspect of the invention.

Further aspects and embodiments of the present invention will be evident from the discussion that follows below.

DETAILED DESCRIPTION

30 As summarised above, the present invention is based, in part, on the recognition that transition metal complexes, for example, of manganese and iron, comprising diazacycloalkane ligands, or chelants, of formula (I) catalyse faster curing of oxidatively curable alkyd-based resin formulations than would have been expected from the prior art.

35 The oxidatively curable resin of the formulation is alkyd-based. As noted above, alkyd resins are a well-understood binder class used in film-forming coating compositions. The term coating composition is to be interpreted broadly and

embraces, for example, varnishes, primary coats, filling pastes and glazes. Coating compositions may be solvent-based or water based, e.g. emulsions. Typical coating compositions comprise solvent-based air-drying coatings and/or paints for domestic use. According to particular embodiments of the present invention, the formulations of the invention (including the fully formulated oxidatively curable coating compositions described herein) are paints. The formulations of the invention (including the fully formulated oxidatively curable coating compositions described herein) may comprise inks, for example a metal plate ink, lithographic ink, relief printing ink, screen ink or offset overprinting ink.

By oxidatively curable alkyd-based resin formulations is meant herein liquids that form a continuous solid coating as a consequence of the course of oxidative reactions (curing) and generally evaporation of a liquid continuous phase (generally solvent).

Typically, curing results in formation of cross-linkages and other bond formations through reactions involving unsaturated components within alkyd-based resin formulations.

In alkyd-based resin formulations, also referred to herein as alkyd-based formulations, the major binder present is an alkyd. By binder is meant in the art and herein the film-forming (curable) component within curable compositions, i.e. the component within the compositions that forms the desired three-dimensional network upon curing.

Typically the curable component of an oxidatively curable composition (e.g. a formulation of the invention) will comprise between about 1 and about 98 % by weight, for example between about 1 and about 90% by weight of the total weight of the composition, e.g. between about 20 and about 70% by weight of the total weight of the composition. At least 50% by weight of the oxidatively curable portion (i.e. of the binder) in an oxidatively curable alkyd-based resin, i.e. from about 50% by weight to about 100% by weight, is curable alkyd resin. Typically, at least 75% by weight of the binder in an oxidatively curable alkyd-based resin, i.e. from about 75% by weight to about 100% by weight (e.g. from about 90% by weight to about 100% by weight), is curable alkyd resin. According to particular embodiments, about 100% by weight of the binder in an oxidatively curable alkyd-based resin is curable alkyd resin. The balance, if any, of the curable (i.e. binder) component may be, for example, curable acrylate, urethane, polybutadiene and epoxy ester resins. The skilled person is aware that introducing quantities of curable binders other than curable alkyds allows the distinct properties of such binders to be introduced to a controllable degree into the ultimate

coating resultant from application of a composition, such as an oxidatively curable composition, which may be made from the formulation of the invention.

As described above, oxidatively curable alkyd resins are a well-understood and indeed dominant binder in many oxidatively curable paints (both for commercial and domestic use) and other coating compositions. They are employed, in particular, in solvent-based coating compositions.

Alkyds (used synonymously herein with alkyd resins) are produced by the condensation, typically polycondensation, of polyols with carboxylic acids or anhydrides. To make them susceptible to the so-called drying process, some alkyd resins (i.e. those that are oxidatively curable, present in the formulation of the invention) are reacted with unsaturated triglycerides or other source of unsaturation. Plant and vegetable oils, such as linseed oil, are frequently used as the source of triglycerides. The term oxidatively curable alkyd resin thus generally refers in the art, and herein, to polyesters modified with fatty acids. As is known in the art, alkyd resins are generally prepared via condensation polymerisation reactions between three types of monomers: (i) one or more polyalcohols (also known as polyols), (ii) one or more polybasic acids (also known as polyacids); and (iii) long chain unsaturated fatty acids or triglyceride oils, which confer upon the alkyds the susceptibility towards curing.

Owing to its presence in naturally occurring oils, glycerol is a widely used polyol in the preparation of alkyds. Other examples of suitable polyhydric alcohols include: pentaerythritol, dipentaerythritol, ethylene glycol, diethylene glycol, propylene glycol, neopentyl glycol, trimethylol propane, trimethylol ethane, di-trimethylol propane and 1,6-hexane diol.

Polycarboxylic acids and the corresponding anhydrides, used to synthesise alkyds, comprise aromatic, aliphatic and cycloaliphatic components, which are generally derived from petrochemical feedstocks. Typical examples of such polyacids include: phthalic acid and its regioisomeric analogues, trimellitic acid, pyromellitic acid, pimelic acid, adipic acid, azelaic acid, sebacic acid, maleic acid, fumaric acid and tetrahydrophthalic acid.

Suitable so-called drying and semi-drying fatty acids or mixture thereof, useful herein, are typically ethylenically unsaturated conjugated or non-conjugated C_{22} - C_{24} carboxylic acids, such as oleic, ricinoleic, linoleic, linolenic, licanic acid and eleostearic acids or mixture thereof, typically used in the forms of mixtures of fatty acids derived from natural or synthetic oils.

By semi-drying and drying fatty acids is meant fatty acids that have the same fatty acid composition as the oils (i.e. the esters) from which they are derived. The classification of the oils is based on the iodine number: for a drying oil the iodine

number is >140; for a semi-drying oil the iodine number is ranging between 125 and 140, and for a non-drying oil the iodine number is <125 (see "*Surface Coatings*", part 1, Chapman & Hall, London, page 55, 1993).

Typically, oxidatively curable alkyd-based formulations, both generally and
5 according to the first aspect of the invention, are liquids. More typically still, such formulations are solvent-based, that is to say they comprise an organic solvent (which may be a mixture of solvents) for the binder and, in accordance with the first aspect of the invention, the chelant.

In other words, "solvent-based" implies to the skilled person in this context
10 formulations that are based on organic (i.e. non-aqueous) solvents, i.e. comprising an organic solvent as a liquid continuous phase. Examples of suitable solvents include aliphatic (including alicyclic and branched) hydrocarbons, such as hexane, heptane, octane, cyclohexane, cycloheptane and isoparaffins; aromatic hydrocarbons such as toluene and xylene; ketones, e.g. methyl ethyl ketone and methyl isobutyl ketone;
15 alcohols, such as secondary butanol, isopropyl alcohol, n-butyl alcohol and n-propyl alcohol, glycols such as propylene glycol; alcohol ethers and esters, glycol monoethers, such as the monoethers of ethylene glycol and diethylene glycol; monoether glycol acetates, such as 2-ethoxyethyl acetate; N-methylpyrrolidone; as well as mixtures thereof. Isomeric variants are included. Thus, for example, the term
20 hexane embraces mixtures of hexanes. According to particular embodiments of the invention, the solvent is a hydrocarbyl (i.e. hydrocarbon) solvent, e.g. an aliphatic hydrocarbyl solvent, e.g. solvents comprising mixtures of hydrocarbons. Examples include white spirit and solvents available under the trademarks Shellsol, from Shell Chemicals and Solvesso and Exxsol, from Exxon.

25 Whilst according to many embodiments of the various aspects of the present invention compositions and formulations are solvent-based, water-based alkyd-based resin formulations and coating compositions are also well known and the compositions and formulations described herein may be water-based (i.e. comprise water as a continuous liquid phase). Accordingly, compositions and formulations described herein
30 may be of alkyd-based resin formulations in the form of emulsions, and may thus comprise a suitable emulsifier, as is well known in the art.

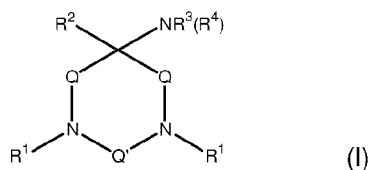
When an alkyd-based formulation or composition is referred to herein as "oxidatively curable", it is to be understood that this term is being used to describe a composition susceptible to the reactions that occur between unsaturated groups (e.g.
35 carbon-carbon double bonds) and oxygen from the air, which reactions constitute oxidative curing and are manifested in hardening and formation of solid coatings obtainable from such compositions or formulations. Thus, an oxidatively curable alkyd-

based resin formulation is a formulation capable of oxidative curing, but which has not yet been allowed to cure. In contrast, the composition of the third aspect of the invention is directed towards formulations after curing, i.e. when cured. The formation of the desired coating resultant from curing may be accelerated through the use of catalytic drying, for example by transition metal-based driers, in particular transition metal-based driers comprising chelants of formula (I).

A characteristic feature of the various aspects of the present invention is the use of chelants of formula (I). When coordinated to suitable transition metal ions, the resultant complexes accelerate the curing of the oxidatively curable formulation of the invention, which acceleration is absent in the absence of suitable transition metal ions.

The nature of the chelants of formula (I) will now be described. It will be understood that more than one such chelant may be used in accordance with the various aspects of the invention. Typically, however, only one kind of chelant will be used.

The chelant is of formula (I):



wherein:

each of $-R^1$, $-R^2$, $-R^3$ and $-R^4$ independently represents $-H$, $-C_{1-24}$ alkyl, $-C_{6-10}$ aryl or a group comprising a heteroatom capable of coordinating to a metal ion;

Q represents methylene or ethylene, in which one or more hydrogen atoms may be optionally independently replaced with a C_{1-24} alkyl or a C_{6-10} aryl group; and

Q' represents ethylene or n-propylene, in which one or more hydrogen atoms may be optionally independently replaced with a C_{1-24} alkyl or a C_{6-10} aryl group.

According to particular embodiments, where any of R^1 , R^2 , R^3 or R^4 is a C_{1-24} alkyl, this may be a C_{1-10} alkyl, which according to still more particular embodiments may be a C_{1-6} alkyl, e.g. methyl.

Where any of R^1 , R^2 , R^3 or R^4 is a group comprising a heteroatom capable of coordinating to a metal ion, such groups may be the same or different. The heteroatom is typically present in a heteroaryl or non-aromatic heterocyclic ring, often a heteroaryl-containing group, which is optionally substituted with one or more (typically no or one) C_{1-4} alkyl groups. In particular embodiments, groups containing a heteroatom comprise one or more nitrogen atoms, for example one or two nitrogen atoms, often one nitrogen atom; and/or the ring containing the heteroatom (e.g. the one or more nitrogen atoms,

for example one or two nitrogen atoms, often one nitrogen atom) is connected to the remainder of formula (I) through an alkylene linker, typically a straight-chain alkylene linker, comprising from 1 to 6 carbon atoms (i.e. typically methylene, ethylene, n-propylene, n-butylene, n-pentylene and n-hexylene), often methylene or ethylene and particularly often methylene.

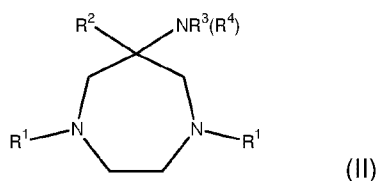
Where one or more of R^1 , R^2 , R^3 and R^4 comprises a heteroaryl group as described herein, the heteroaryl group may be, for example, pyridine, pyrimidine, pyrazine, pyrazole, imidazole, benzimidazole, quinoline, quinoxaline, 1,2,3- or 1,2,4-triazole, isoquinoline, carbazole, indole, isoindole, oxazole and thiazole. According to particular embodiments, the heteroaryl group is pyridine. Where any of R^1 , R^2 , R^3 or R^4 comprises a heteroaryl group, the heteroaryl group may be optionally substituted one or more times with C_{1-4} alkyl groups. Typically, any heteroaryl groups in R^1 , R^2 , R^3 or R^4 are either unsubstituted or substituted once with a C_{1-4} alkyl group. In particular embodiments, such heteroaryl groups are unsubstituted.

Typically, although not necessarily, where one or more of R^1 , R^2 , R^3 and R^4 comprises a pyridine ring, this is connected to the remainder of formula (I) through the 2-position (i.e. the heteroaryl group is an optionally C_{1-4} alkyl-substituted 2-pyridyl, e.g. 2-pyridyl). More typically still, although also not necessarily, pyridyl groups (in particular 2-pyridyl groups) are connected to the remainder of formula (I) through alkylene linkers (as described herein), for example methylene. According to particular embodiments, one or more of R^1 , R^2 , R^3 and R^4 is 2-pyridylmethyl. According to other particular embodiments one or more of R^1 , R^2 , R^3 and R^4 is 2-pyridylmethyl and optionally R^2 is hydrogen or methyl.

Q and Q' are optionally substituted alkylene groups as hereinbefore defined. The C_{1-24} alkyl groups with which these alkylene groups may be substituted are typically C_{1-18} alkyl groups. The C_{6-10} aryl groups may be phenyl or naphthyl. According to other particular embodiments, Q is an optionally substituted methylene and Q' is an optionally substituted ethylene. According to still more particular embodiments, Q and Q' are unsubstituted, for example Q is an unsubstituted methylene and Q' is an unsubstituted ethylene.

Reflecting, in part, some of the particular embodiments of chelants of formula (I) described above, particular embodiments of chelants of formula (I) may be defined by chelants of formula (II):

12



wherein:

each -R¹ is independently -H, -C₁₋₂₄alkyl, -C₆₋₁₀aryl or pyridine-2ylmethyl, which aryl or pyridinyl is optionally substituted with C₁₋₄alkyl;

5 -R² represents -H or -CH₃; and

each -R³ and -R⁴ is independently -C₁₋₂₄alkyl, -C₆₋₁₀aryl or pyridine-2ylmethyl, which aryl or pyridinyl is optionally substituted with C₁₋₄alkyl,

for example wherein:

10 each -R¹ is independently -H, -C₁₋₂₄alkyl or pyridine-2ylmethyl, which pyridinyl is optionally substituted with C₁₋₄alkyl;

-R² represents -H or -CH₃; and

each -R³ and -R⁴ is independently -C₁₋₂₄alkyl or pyridine-2ylmethyl, which pyridinyl is optionally substituted with C₁₋₄alkyl.

15 In many embodiments of the chelants of formula (II) (and (I)), the two R¹ groups are the same.

According to further particular embodiments of the chelants of formula (II) (and (I)), each R¹ independently represents methyl or pyridine-2ylmethyl, -R² represents methyl and -R³ and -R⁴ each independently represents a -C₁₋₂₄alkyl or -C₆₋₁₀aryl or pyridine-2ylmethyl.

20 Particular chelants of formulae (I) and (II) are:

6-dimethylamino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;

6-amino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;

1,4,6-trimethyl-6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-diazacycloheptane;

6-amino-1,4,6-trimethyl-1,4-diazacycloheptane;

25 6-dimethylamino-1,4,6-trimethyl-1,4-diazacycloheptane;

1,4,6-trimethyl-6-(pyridin-2-ylmethyl)amino)-1,4-diazacycloheptane;

6-{N,N-bis(pyridin-2-ylmethyl)amino}-1,4,6-trimethyl-1,4-diazacycloheptane; and

6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane.

30 The chelant of formula (I) is typically present in formulations according to the present invention in concentrations of from 0.00005 to 0.5 % by weight, often from 0.0001 to 0.1 % by weight.

Where percentages by weight are referred to herein (wt % or % w/w), these mean, unless a context clearly dictates to the contrary, percentages by weight with respect to the binder component (i.e. the alkyd-based resin and any other binders present). With an oxidatively curable alkyd-based coating formulation, for example, the combined weights of the binders are those with respect to which weight percentages herein are based. For example, where a formulation according to the first aspect of the invention comprises 0.00005 % w/w of chelant of formula (I), this is with respect to the weight of the curable components of the composition (i.e. the weight of the binder(s)).

Often, formulations of the first aspect of the invention will comprise a complex of the chelant of formula (I) with a suitable transition metal ion. However, this need not necessarily be so. This is because we have recognised that there is technical advantageousness in providing an oxidatively curable alkyd-based resin formulation comprising a chelant (used interchangeably herein with the term "chelating agent") of formula (I), which formulation is essentially absent at least manganese, iron, cobalt, vanadium and copper ions. These, ions, if present in the formulation, can form together with the chelant a metal drier capable of accelerating oxidative curing.

A manufacturer of an alkyd-based resin formulation suitable for oxidative curing can thus include a chelant of formula (I) in an amount appropriate for a given oxidatively curable alkyd-based resin formulation. Each type of oxidatively curable alkyd-based resin can, and typically does, have different sensitivity towards radical curing and may thus require a particular concentration of a metal drier for optimal curing. However, to determine the appropriate concentration in practice is not straightforward, since a metal drier, for example a manganese- or iron-based catalyst, can initiate radical curing before the coating composition (e.g. paint) comprising an oxidatively curable alkyd-based resin formulation (and other components) can be applied, leading to undesirable degradation and/or hardening of the resin formulation. In contrast, a manufacturer of an oxidatively curable alkyd-based resin formulation, as opposed to the manufacture of a fully formulated oxidatively curable coating composition comprising such an oxidatively curable alkyd-based resin formulation, can determine the optimum amount of metal drier for a given alkyd-based resin formulation and add to batches of it a suitable amount of chelant of formula (I) (but not the transition metal ions that allow formation of a catalytically active drier, which are often, but not necessarily, manganese, iron, cobalt, vanadium or copper ions). An appropriate quantity of transition metal ions can then be added to the resultant formulation by, for example, a manufacturer of a coating composition, along with other components to make a fully formulated oxidatively curable coating composition.

Mixing of appropriate chelants of formula (I) with alkyd-based resin formulations in the essential absence of at least manganese, iron, cobalt, vanadium and copper ions which, if present, render these chelants catalytically active as metal driers, affords formulations at least less susceptible to skinning or instability without the requirement to add antiskinning agents and/or to take other specific measures to avoid skinning. Such formulations thus constitute particular embodiments of the first aspect of the invention.

The embodiments of the formulation of the first aspect of the invention that are essentially absent at least manganese, iron, cobalt, vanadium and copper ions comprise less than 0.001 % by weight of at least ions of each of manganese, iron, cobalt, vanadium and copper. By this is meant that the formulation of the invention is absent 0.001% by weight manganese ions, absent 0.001% by weight iron ions, absent 0.001% by weight cobalt ions, absent 0.001% by weight vanadium ions and absent 0.001% by weight copper ions. An appropriate quantity of suitable transition metal cations (e.g. ions of one or more of manganese, iron, cobalt, vanadium and copper) can be added after preparation of such a formulation, for example when introducing optional additional components to form an oxidatively curable coating composition.

Particular formulations of the invention can if desired comprise less than 0.0001% by weight of each of at least manganese, iron, cobalt, vanadium and copper ions. Still other embodiments of formulations of the invention comprise less than 0.001 % by weight of each of manganese, iron, cobalt, vanadium, copper, titanium, molybdenum, tungsten, cerium and zirconium. Still other embodiments of formulations of the invention comprise less than 0.0001 % by weight of each of manganese, iron, cobalt, vanadium, copper, titanium, molybdenum, tungsten, cerium and zirconium.

Ideally, embodiments of the formulation of the invention with concentrations of specific transition metal ions of less than 0.001 wt% or 0.0001 wt% are absent any of the transition metal ions specified. Obviously, however, this is in practice impossible to achieve. Accordingly, these formulations are preferably absent of the transition metal ions specified to the greatest extent practicable.

In order to make, in accordance with a method of the second aspect of the invention, a formulation according to the first aspect of the invention, a composition comprising an oxidatively curable alkyd-based resin is contacted with a composition comprising a chelant of formula (I). The composition comprising the chelant of formula (I) that is contacted with the composition comprising the alkyd-based resin may, in some embodiments, comprise a transition metal ion-containing complex comprising the chelant. In other embodiments, the composition comprising the chelant of formula (I) is not part of a transition metal ion-containing complex, in which case a source of

transition metal ions may, if wished, be added afterwards (or indeed have been formulated together with the alkyd-based resin before addition of the chelant) so as to form a complex comprising the chelant of formula (I) *in situ* (i.e. within the alkyd-based resin formulation). Both of these different types of embodiments are described below.

5 The typical molar ratio between any transition metal ions and the chelant is between about 0.1:1 and about 10:1, often between about 0.3:1 and about 3:1. Often, the molar ratio between chelant and transition metal ions will be approximately 1:1. However, this need not necessarily be the case. Without being bound to theory, an excess of transition metal ions may be beneficial to allow some adsorption on solid particles
10 without losing too much siccative activity. On the other hand, a stoichiometric excess of chelant may be beneficial to improve regeneration of catalytically active species during curing, which can lead to improved drying (i.e. curing) performance despite using a lower quantity of transition metal ions. Using a stoichiometric excess of chelant can also be advantageous by reducing the intensity of coloured metal complexes. The skilled person
15 will be able to take into account these considerations when preparing oxidatively curable coating compositions, for example formulations of the invention.

 The contacting of the method of the first aspect of the invention may be during formulation of fully formulated oxidatively curable alkyd-based resin coating compositions (described below), particularly if the chelant of formula (I) is part of the
20 complex comprising a suitable transition metal ion.

 If the chelant of formula (I) introduced as a transition metal ion-containing complex, the complex may be prepared by contacting the chelant of formula (I) with a suitable transition metal salt in a suitable solvent, by which is meant that either or both of the chelant and transition metal salt may be in a suitable solvent prior to contact with
25 each other. The salt can be a soap. The resultant complexes-containing mixture may then be contacted with a composition comprising an oxidatively curable alkyd-based resin, which is typically dissolved in an organic solvent described above when describing solvent-based alkyd-based formulations (or emulsified in a water-based liquid such as those described above when describing water-based alkyd-based formulations).

30 Often, the metal salt used will be a manganese salt or iron salt, typically of a divalent or trivalent redox state. Upon contacting the manganese or iron (or other transition metal) salt with the chelant, formation of manganese- or iron-chelant complexes (or other transition metal-chelant complexes) takes place.

 The transition metal salt used can be a solid, a suspension, or as a solution in a
35 variety of solvents. Typically the salt comprises a manganese (II), manganese (III),

iron (II), or iron (III) ion, although other salts, e.g. manganese (IV) salts may also be used. Such manganese or iron (or other metal ion) salts can be added as solids, suspensions, or as solutions in a variety of solvents. The invention contemplates use of a mixture of metal salts although a single salt is typically used.

5 Adding the chelant as a solution can be advantageous in permitting improved and/or easier mixing with the (solution of) binder(s). It may be beneficial to dilute chelant in a suitable solvent before adding to the binder if it is wished to introduce a very small amount of chelant, so greater accuracy of dosing can be achieved. Depending on the properties of the chelant and the desired resin-chelant formulation,
10 suitable solvents include aliphatic hydrocarbons, such as heptanes, water, alcohols, such as ethanol or propylene glycol, or mixtures thereof. The skilled person will be able to easily formulate such solutions, generally using solvent such as those described above.

It will be understood that there is no particular limitation as to the source of the
15 transition metal ions. Typically, however, salts are selected from the group consisting of optionally hydrated MnCl_2 , FeCl_2 , FeCl_3 , MnBr_2 , $\text{Mn}(\text{NO}_3)_2$, $\text{Fe}(\text{NO}_3)_3$, MnSO_4 , FeSO_4 , $(\text{Fe})_2(\text{SO}_4)_3$, $\text{Mn}(\text{acetylacetonate})_2$, $\text{Fe}(\text{acetylacetonate})_2$, $\text{Mn}(\text{acetylacetonate})_3$, $\text{Fe}(\text{acetylacetonate})_3$, $\text{Mn}(\text{R}_5\text{COO})_3$ (including $\text{Mn}(\text{acetate})_3$), $\text{Fe}(\text{R}_5\text{COO})_3$, $\text{Mn}(\text{R}_5\text{COO})_2$ and $\text{Fe}(\text{R}_5\text{COO})_2$, wherein R_5 is selected from a C_1 - C_{24} alkyl. Where the
20 salt comprises two or more R_5 groups, these can be the same or different. The alkyl moieties, by which is meant saturated hydrocarbyl radicals, may be straight-chain or comprise branched and/or cyclic portions. Indeed, throughout the specification, where reference is made to alkyl, unless the context dictates to the contrary, this means a C_{1-24} alkyl, which may be straight-chain or branched and may be cycloalkyl or comprise a
25 cyclic portion (e.g. alkyl may be cyclohexylmethyl), for example C_{1-10} alkyl or C_{1-6} alkyl, e.g. methyl.

Often, the manganese or iron salt is selected from $\text{Mn}(\text{R}_5\text{COO})_2$ or $\text{Fe}(\text{R}_5\text{COO})_2$, particularly with $\text{R}_5\text{COO}(\cdot)$ being selected from acetate, octanoate, 2-ethylhexanoate, neodecanoate (3,3,5,5-tetramethylhexanoate), and naphthenate. Very
30 often, a manganese salt is used. Particularly often, a manganese salt is used, which, for example is selected from manganese(octanoate)₂, manganese(naphthenate)₂, manganese(2-ethylhexanoate)₂ and manganese (neodecanoate)₂. The invention also contemplates use of a mixture of different redox states of the metal ions with the same counterion, for example a mixture of manganese(2-ethylhexanoate)₂ and
35 manganese(2-ethylhexanoate)₃.

The term optionally hydrated is well known in the art. Metal salts often contain water molecules within a crystal lattice, which will remain present unless the hydrated

metals salts are subjected to specific drying steps by heating or drying under reduced pressure. However, partially or fully dehydrated metal salts can also be used. For example, manganese (II) acetate, manganese (II) chloride and iron (II) chloride can be bought as tetrahydrate salts or as dehydrated salts. Iron (III) chloride can be purchased as the anhydrous salt as well as the hexahydrate salt. Commercial manganese sulfate is available in both tetrahydrate and monohydrate forms.

Often these transition metal salts are commercially available as solutions, particularly if they are of the formulae $Mn(R_5COO)_2$ or $Fe(R_5COO)_2$ described above, for example in hydrocarbon solutions to facilitate dissolution in the solvent-based curable compositions such as paint formulations. However, other solvents may also be used, including alcohols and water (or aqueous solutions), especially for chloride, sulfate and acetate salts of manganese and iron ions.

Formulations of the invention that comprise less than 0.001 % (or 0.0001 %) by weight of ions of each of at least manganese, iron, cobalt, vanadium and copper may be prepared by contacting chelant of formula (I) with (e.g. adding it to) an oxidatively curable alkyd-based binder, typically dissolved in an organic solvent described above (or emulsified in a water-based liquid), as described above. The chelant may be added as a pure material to the resin(s), or as a solution. Adding the chelant as a solution can be advantageous in permitting improved and/or easier mixing with the (solution of) binder(s). It may be beneficial to dilute chelant in a suitable solvent before adding to the binder if it is wished to introduce a very small amount of chelant, so greater accuracy of dosing can be achieved. Depending on the properties of the chelant and the desired resin-chelant formulation, suitable solvents include aliphatic hydrocarbons, such as heptanes, water, alcohols, such as ethanol or propylene glycol, or mixtures thereof. The skilled person will be able to easily formulate such solutions, generally using a solvent such as those described above.

Thus, as described herein, formulations of the invention comprising transition metal ion-containing complexes of the chelant of formula (I) can be prepared, either by contacting an alkyd-based resin composition with such a complex directly, or by contacting an alkyd-based resin composition with chelant that is not part of such a complex and then adding to the resultant formulation a source of transition metal ions. As a still further embodiment of the method of the second aspect of the invention, an alkyd-based resin composition comprising suitable transition metal ions may be contacted with chelant of formula (I). Generally, formulations of the invention comprising transition metal ions comprise a concentration of between about 0.0003 wt% and about 0.07 wt%, for example about 0.0005 wt% and about 0.05 wt%, e.g.

between about 0.005 wt% and about 0.05 wt%, of the suitable transition metal ion, such as those described immediately below.

Transition metal ions to which the chelants of formula (I) may coordinate, to provide metal driers (transition metal ion-containing complexes that can accelerate curing of the oxidatively curable alkyd-based resin in the formulation of the invention may be, for example manganese, iron, cobalt, vanadium, copper, titanium, molybdenum, tungsten, cerium and zirconium ions, more typically manganese, iron, cobalt, vanadium and copper ions, still more typically manganese or iron ions, or mixtures of any these metal ions. The valency of the metal ions may range from +1 to +6, often from +2 to +5. Examples include metal ions selected from the group consisting of Mn(II), Mn(III), Mn(IV), Mn(V), Cu(I), Cu(II), Cu(III), Fe(II), Fe(III), Fe(IV), Fe(V), Co(I), Co(II), Co(III), Ti(II), Ti(III), Ti(IV), V(II), V(III), V(IV), V(V), Mo(II), Mo(III), Mo(IV), Mo(V), Mo(VI), W(IV), W(V), W(VI), Ce(III), Ce(IV) and Zr(IV), for example metal ions selected from the group consisting of Mn(II), Mn(III), Mn(IV), Mn(V), Cu(I), Cu(II), Cu(III), Fe(II), Fe(III), Fe(IV), Fe(V), Co(I), Co(II), Co(III), V(II), V(III), V(IV) and V(V), for example metal ions selected from the group consisting of Mn(II), Mn(III), Mn(IV), Fe(II) and Fe(III).

In complexes comprising the chelant of formula (I), the number of metal ions per chelant molecule will depend on both the chelant molecule and on the presence of other ligands in the complex, in particular coordinating ligands, including coordinating ligands capable of bridging metal ions. For example, if the chelant of formula (I) contains only three nitrogen donor atoms, the complex containing it will typically contain one metal ion per chelant molecule. Complexes comprising more than one metal ion (and thus more than one chelant molecule) may be formed if metal ions are bridged by, for example, oxide or hydroxyl anions, or carboxylate groups, giving rise to dinuclear or multinuclear complexes. Typically, however, the molar ratio between metal ions and chelants of formula (I) in complexes of them is 1:1.

Complexes comprising the chelant of formula (I) may, for example, be of the generic formula (III):

30



in which:

M represents a metal ion selected from Mn(II), Mn(III), Mn(IV), Mn(V), Cu(I), Cu(II), Cu(III), Fe(II), Fe(III), Fe(IV), Fe(V), Co(I), Co(II), Co(III), Ti(II), Ti(III), Ti(IV), V(II), V(III), V(IV), V(V), Mo(II), Mo(III), Mo(IV), Mo(V), Mo(VI), W(IV), W(V), W(VI), Ce(III), Ce(IV), Zr(IV);

19

each X independently represents a coordinating species selected from any mono-, bi-, or tri-charged anions and any neutral molecule able to coordinate a metal ion M in a mono-, bi- or tridentate manner;

each Y is independently a non-coordinating counterion;

5 a represents an integer from 1 to 10;

k represents an integer from 1 to 10;

n represents an integer from 1 to 10;

m represents an integer from 1 to 20; and

L represents a chelant of formula (I),

10 or a hydrate thereof.

According to particular embodiments of formula (III), alone or in combination:

M represents a metal ion selected from Mn(II), Mn(III), Mn(IV), Fe(II) and Fe(III);

15 X represents a coordinating species selected from O^{2-} , $[R^6BO_2]^{2-}$, R^6COO^- , $[R^6CONR^6]^-$, OH^- , NO_3^- , NO , S^{2-} , R^6S^- , PO_4^{3-} , HPO_4^{2-} , $H_2PO_4^-$, $[PO_3OR^6]^{3-}$, H_2O , CO_3^{2-} , HCO_3^- , R^6OH , $NR^6R^7R^8$, R^6OO^- , O_2^{2-} , O_2^- , R^6CN , Cl^- , Br^- , I^- , OCN^- , SCN^- , CN^- , N_3^- , F^- , RO^- , ClO_4^- , $CF_3SO_3^-$;

Y represents a counterion selected from ClO_4^- , $CF_3SO_3^-$, $[B(R^6)_4]^-$, $[FeCl_4]^-$, PF_6^- , R^6COO^- , NO_3^- , RO^- , $N^+R^6R^7R^8R^9$, Cl^- , Br^- , I^- , F^- , $S_2O_6^{2-}$, OCN^- , SCN^- , H_2O , BF_4^- , SO_4^{2-} ; R^6 , R^7 , R^8 and R^9 each independently represents hydrogen, optionally substituted

20 alkyl or optionally substituted aryl;

a represents an integer from 1 to 4;

k represents an integer from 1 to 10;

n represents an integer from 1 to 4; and

m represents an integer from 1 to 8.

25 By aryl in formula (III) is meant herein a C_{6-14} aryl (e.g. phenyl or naphthyl) and by alkyl is meant a C_{1-24} alkyl.

Where an alkyl or aryl group is optionally substituted, this may be with one or more substituents independently selected from the group consisting of -halo, -OH, -OR¹⁰, -NH₂, -NHR¹⁰, -N(R¹⁰)₂, -N(R¹⁰)₃⁺, -C(O)R¹⁰, -OC(O)R¹⁰, -CO₂H, -CO₂⁻, -CO₂R¹⁰, -C(O)NH₂, -C(O)NHR¹⁰, -C(O)N(R¹⁰)₂, -heteroryl, -R¹⁰, -SR¹⁰, -SH, -P(R¹⁰)₂, -P(O)(R¹⁰)₂, -P(O)(OH)₂, -P(O)(OR¹⁰)₂, -NO₂, -SO₃H, -SO₃⁻, -S(O)₂R¹⁰, -NHC(O)R¹⁰ and -N(R¹⁰)C(O)R¹⁰, wherein each R¹⁰ is independently selected from alkyl, aryl, aralkyl optionally substituted one or two more times with a substituent selected from the group consisting of -halo, -NH₃⁺, -SO₃H, -SO₃⁻, -CO₂H, -CO₂⁻, -P(O)(OH)₂, -P(O)(O⁻)₂.

35 According to particular embodiments, a=k=1.

As is known, the ability of metal driers to catalyse the curing of oxidatively curable coating compositions arises from their ability to participate in redox chemistry: the nature of the counterion(s) Y are not of great importance. The choice of these may be affected by the solubility of the complex of metal ions and chelant of formula (I) in a given formulation or composition. For example, counterion(s) Y such as chloride, sulfate or acetate may serve to provide a readily water-soluble complex, if a water-based paint is desired. When using solvent-based (i.e. non-aqueous) compositions, it may be desirable to use larger, less popular counterions such as 2-ethylhexanoate. Suitable counterion(s) Y (and coordinating species X) may be selected without difficulty by the skilled person.

According to particular embodiments, X and Y may be independently selected from the group consisting of bromide, iodide, nitrate, sulfate, methoxide, ethoxide, formate, acetate, propionate, 2-ethylhexanoate, naphthenate, and hydroxide.

An example of a neutral molecule able to coordinate the metal in a mono-, bi- or tridentate manner is acetonitrile, for example, to afford a complex of the formula $[ML(CH_3CN)]Cl_2$.

It will be understood that counterions Y serve to balance the charge resultant from the complex formed by the metal ion(s) M, coordinating species X and chelant(s) L. Thus, if the charge on the complex is positive, there will be one or more anions Y. Conversely, there will be one or more cations Y if the charge on the complex is negative.

Particular examples of complexes comprising the chelant L of formula (I) include $[MLCl_2]$, $[MLCl]Cl$, $[ML(H_2O)](PF_6)_2$, $[ML]Cl_2$, $[MLCl]PF_6$ and $[ML(H_2O)](BF_4)_2$. In particular embodiments, transition metal ion M in this context is an ion of manganese or iron.

It will be understood from the foregoing discussion that complexes of formula (III) embrace dinuclear complexes (i.e. comprising two metal ions M), such as those containing hydroxide, oxo, carboxylate or halide as bridging ligands (with a bridging ligand indicated with μ). Depending on the denticity of the chelant L (i.e. the number of atoms through which it chelates), one or more bridging molecules may be present. A combination of bridging and non-bridging ligands X may be present. Non-limiting examples of dinuclear complexes include $[LM(\mu-RCOO)_2ML](RCOO)_2$ (wherein RCOO = methyl (acetate) or 2-ethylhexanoate with L being a tetradentate chelant, M = Mn(II) or Fe(II); and $LM(\mu-O)(\mu-RCOO)_2ML](RCOO)_2$ with L being a tridentate chelant, M = Mn(III) or Fe(III). The metal-ligand complex, containing suitable counterion(s) Y, can

be contacted with (e.g. added to) an alkyd-based resin so as to form a formulation of the present invention.

A formulation of the invention can, and generally will, be used in the manufacture of a fully formulated oxidatively curable coating composition. By the term
5 "fully formulated oxidatively curable coating composition" is implied, as is known to those of skill in the art, oxidatively curable formulations that comprise additional components over and above the binder (the oxidatively curable material, which is predominantly oxidatively curable alkyd resin according to the present invention), an aqueous or non-aqueous solvent/liquid continuous phase and any metal driers
10 intended to accelerate the curing process. Such additional components are generally included so as to confer desirable properties upon the coating composition, such as colour or other visual characteristics such as glossiness or mattness, physical, chemical and even biological stability (enhanced biological stability being conferred upon coating compositions by the use of biocides for example), or modified texture,
15 plasticity, adhesion and viscosity.

For example, such optional additional components may be selected from solvents, antioxidants (sometimes referred to as antiskinning agents), additional siccatives (i.e. not comprising chelants of formula (I)), auxiliary driers, colourants (including inks and coloured pigments), fillers, plasticisers, viscosity modifiers, UV light
20 absorbers, stabilisers, antistatic agents, flame retardants, lubricants, emulsifiers (in particular where an oxidatively curable coating composition or formulation of the invention is aqueous-based), anti-foaming agents, viscosity modifiers, antifouling agents, biocides (e.g. bactericides, fungicides, algacides and insecticides), anticorrosion agents, antireflective agents, anti-freezing agents, waxes and thickeners.
25 Typically, formulations prepared in accordance with embodiments of the method of the second aspect of the invention will comprise at least an organic solvent, selected from the list of solvents described above and a filler, and generally an antiskinning agent, in addition to the alkyd and optionally other binders and chelant present in the formulation of the invention. The skilled person is familiar with the incorporation of these and other
30 components into oxidatively curable coating composition so as to optimise such compositions' properties.

It will be appreciated that some of these optional additional components possess more than one functional property. For example, some fillers may also function as colourants. The nature of any additional components and the amounts
35 used may be determined in accordance with the knowledge of those of skill in the art and will depend on the application for which the curable coating compositions intended. Examples are provided below but these are intended to be illustrative, not limitative.

When producing a fully formulated oxidatively curable coating composition that is, for example, a paint, one or more antioxidants (customarily referred to in the art as antiskinning agents) are often included to avoid premature curing of the oxidatively curable coating composition prior to its use. Such premature curing may be manifested by, for example, the formation of a skin on or lumpy matter in the oxidatively curable coating composition as a result of curing during storage, for example hardening of the surface of a paint layer in a can), owing to the activity of the siccative with oxygen on the oxidatively curable binder. Antiskinning agents are understood to reduce skinning by quenching radicals formed and/or by inactivation of drier catalysts by binding to one or more of the coordination sites. Examples include, but are not limited to, methylethylketoxime, acetoxime, butyraldoxime, methyl-isobutylketoxime, 2-cyclohexylphenol, 4-cyclohexylphenol, t-butyl-hydroquinone, dialkylhydroxylamine, acetylacetonate, ammonia, vitamin E (tocopherol), hydroxylamine, triethylamine, dimethylethanolamine, 2-t-butyl-4-methylphenol, and 2-[(1-methylpropyl)amino]ethanol. According to particular embodiments, the antiskinning agent is selected from the group consisting of methylethylketone-oxime, acetoxime, butyraldoxime, dialkylhydroxylamine, ammonia, hydroxylamine, triethylamine, dimethylethanolamine, o-cyclohexylphenol, p-cyclohexylphenol and 2-t-butyl-4-methylphenol.

The quantity of antiskinning agent present in an oxidatively curable coating composition is typically between about 0.001 and about 2.5 wt%. The antioxidant or antiskinning agent may be added to an alkyd-based resin formulation, e.g. of the invention, together with (or separately from) the chelant prior to or during the preparation of a fully formulated oxidatively curable coating composition (for example a paint or other coating composition).

Colourants include pigments and inks. Titanium dioxide is a pigment commonly included in many coating compositions, in particular paints.

Fillers may be added to an oxidatively curable coating composition for a number of reasons, for example to bulk out the coating composition and to compare particular properties on the cured composition. Typically, fillers will be inorganic solids that are generally introduced in particulate (finely divided) form. Examples include silica, silicates or clays (for example mica, talc, kaolin), carbonate or other minerals and metal salts or oxides (such as marble, quartzite). Other suitable fillers will be evident to the skilled person.

It may be advantageous if an alkyd resin manufacturer has determined a particular concentration of metal drier that is appropriate for a particular alkyd-based resin formulation for the manufacturer to recommend to users of the formulation an

appropriate source of transition metal ions that may be added in order to generate a desired metal drier *in situ*.

Moreover, according to the fifth aspect of the invention, there is provided a kit comprising a formulation of the invention comprising less than 0.001% by weight of ions of each of at least manganese, iron, cobalt, vanadium and copper and, separately, a composition comprising a source of suitable transition metal ions, typically ions selected from the group consisting of manganese, iron, cobalt, vanadium and copper ions, more typically either manganese ions or iron ions, often in the form of a salt such as those described above, for admixture to a formulation of the invention. The kit may optionally comprise instructions or other guidance as to methods according to which the formulation and the transition metal ions may be contacted. In this way, the manufacture of a formulation of the invention can, after optimising the nature of the source of transition metal ions, for example by the preparation of a particular solution of a particular transition metal ions salt, optimise the manner in which formulations containing transition metal complexes can be prepared. The preparation of an oxidatively curable alkyd-based coating composition may be by the manufacturer of such compositions (e.g. a paint manufacturer) or by an end consumer of oxidatively curable alkyd-based coating compositions, who can contact a source of transition metal ions with an otherwise fully formulated oxidatively curable alkyd-based coating composition.

It is also within the scope of the current invention that a paint manufacturer, for example, would add commercial metal-soap/chelant mixtures, such as the non-limiting example of Borchers® Dry 0410 (a mixture of bpy with Mn(neodecanoate)₂ commercially available from OMG), as a mixture of bpy with Mn(neodecanoate)₂. The additional chelant present in the alkyd resin will improve the drying behaviour without causing excessive yellowing which may be occurring if more of the Mn-soap/ligand mixture is added to the paint formulation.

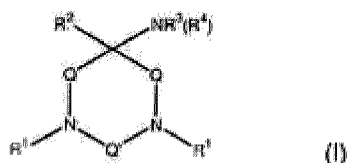
Additionally, one or more auxiliary driers may be added to the fully formulated oxidatively curable coating composition. Such auxiliary driers may be optional additional components within, but are often not present in, the formulation of the invention. Such auxiliary driers include fatty acid soaps of zirconium, bismuth, barium, cerium, calcium, lithium, strontium, and zinc. Typically, fatty acid soaps are optionally substituted octanoates, hexanoates and naphthenates. Without being bound by theory, auxiliary driers (sometimes referred to as through driers) are generally understood to diminish the effect of adsorption of the main drier on solid particles often present in an oxidatively curable coating composition. Other non-metal based auxiliary driers may also be present if desired. These may include, for example, thiol

compounds, as described in US 2001/0008932 A1 (Bakkeren *et al.*) or biomolecules as described in US 2005/0245639 A1 (Oostveen *et al.*) Concentrations of auxiliary driers within oxidatively curable coating compositions (or formulations of the invention) are typically between about 0.01 wt % and 2.5 wt % as is known in the art.

The formulations of the invention (including the fully formulated oxidatively curable coating compositions described herein) may be used as a decorative coating, e.g. applied to wood substrates, such as door or window frames, or for other substrates such as those made of synthetic materials (such as plastics including elastomeric materials), concrete, leather, textile, glass, ceramic or metal, in accordance with the sixth aspect of the invention. The thus-applied composition may then be allowed to cure. In this respect, the third aspect of the invention is directed towards a formulation according to the first aspect, or obtainable according to the second aspect, when cured.

The invention may be further understood with reference to the following non-limiting clauses:

1. A formulation comprising an oxidatively curable alkyd-based resin and a chelant of the formula (I):



wherein:

each of -R¹, -R², -R³ and -R⁴ independently represents -H, -C₁₋₂₄ alkyl, -C₆₋₁₀ aryl or a group comprising a heteroatom capable of coordinating to a metal ion;

Q represents methylene or ethylene, in which one or more hydrogen atoms may be optionally independently replaced with a C₁₋₂₄ alkyl or a C₆₋₁₀ aryl group; and

Q' represents ethylene or n-propylene, in which one or more hydrogen atoms may be optionally independently replaced with a C₁₋₂₄ alkyl or a C₆₋₁₀ aryl group.

2. The formulation of clause 1, wherein Q represents methylene or ethylene and Q' represents ethylene or n-propylene.

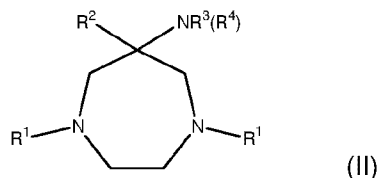
3. The formulation of clause 1 or clause 2, wherein each of -R¹, -R², -R³ and -R⁴ independently represents -H, -C₁₋₁₀alkyl, -C₆₋₁₀aryl or a group comprising a heteroatom capable of coordinating to a metal ion.
- 5 4. The formulation of clause 1 or clause 2, wherein each of -R¹, -R², -R³ and -R⁴ independently represents -H, -C₁₋₆alkyl, -C₆₋₁₀aryl or a group comprising a heteroatom capable of coordinating to a metal ion.
- 10 5. The formulation of clause 1 or clause 2, wherein each of -R¹, -R², -R³ and -R⁴ independently represents -H, -methyl, -C₆₋₁₀aryl or a group comprising a heteroatom capable of coordinating to a metal ion.
- 15 6. The formulation of any one of clauses 1 to 5, wherein the heteroatom capable of coordinating to a metal ion is contained in a heteroaryl or non-aromatic heterocyclic ring, which ring is optionally substituted with a C₁₋₄ alkyl group.
7. The formulation of any one of clauses 1 to 6, wherein the heteroatom capable of coordinating to a metal ion is contained in a heteroaryl ring.
- 20 8. The formulation of clause 6 or clause 7, wherein the heteroaryl ring is unsubstituted.
9. The formulation of any one of clauses 6 to 8, wherein the heteroaryl ring is selected from the group consisting of pyridine, pyrimidine, pyrazine, pyrazole, imidazole, benzimidazole, quinoline, quinoxaline, 1,2,3- or 1,2,4-triazole, isoquinoline, carbazole, indole, isoindole, oxazole and thiazole.
- 25 10. The formulation of clause 9, wherein the heteroaryl ring is pyridine.
- 30 11. The formulation of clause 10, wherein the pyridine ring is connected to the remainder of formula (I) through the 2-position.
12. The formulation of any one of clauses 6 to 11, wherein the ring is connected to the remainder of formula (I) through an alkylene linker.
- 35 13. The formulation of clause 12, wherein the alkylene linker is methylene.

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14. The formulation of any one of clauses 1 to 13, wherein one or more of $-R^1$, $-R^2$, $-R^3$ and $-R^4$ is 2-pyridylmethylene.

15. The formulation of clause 1, wherein the chelant is of formula (II):

5



wherein:

each $-R^1$ is independently $-H$, $-C_{1-24}$ alkyl, $-C_{6-10}$ aryl or pyridine-2ylmethyl, which aryl or pyridinyl is optionally substituted with C_{1-4} alkyl;

10 $-R^2$ represents $-H$ or $-CH_3$; and

each $-R^3$ and $-R^4$ is independently $-C_{1-24}$ alkyl, $-C_{6-10}$ aryl or pyridine-2ylmethyl, which aryl or pyridinyl is optionally substituted with C_{1-4} alkyl.

16. The formulation of clause 1, wherein the chelant is selected from the group consisting of:

15

6-dimethylamino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;

6-amino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;

1,4,6-trimethyl-6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-diazacycloheptane;

6-amino-1,4,6-trimethyl-1,4-diazacycloheptane;

20

6-dimethylamino-1,4,6-trimethyl-1,4-diazacycloheptane;

1,4,6-trimethyl-6-(pyridin-2-ylmethyl)amino)-1,4-diazacycloheptane;

6-{N,N-bis(pyridin-2-ylmethyl)amino}-1,4,6-trimethyl-1,4-diazacycloheptane; and

6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane

25

17. The formulation of any one of clauses 1 to 16, wherein the chelant is present in the formulation at a concentration of between about 0.00005 and about 0.5 wt% with respect to curable resin.

30

18. The formulation of any one of clauses 1 to 16, wherein the chelant is present in the formulation at a concentration of between about 0.001 and about 0.1 wt% with respect to curable resin.

19. The formulation of any one of clauses 1 to 18, which is a solvent-based formulation.
20. The formulation of any one of clauses 1 to 19, further comprising an antiskinning agent.
21. The formulation of clause 20, wherein the antiskinning agent is selected from the group consisting of methylethylketone-oxime, acetoxime, butyraldoxime, dialkylhydroxylamine, ammonia, hydroxylamine, triethylamine, dimethylethanolamine, o-cyclohexylphenol, p-cyclohexylphenol and 2-t-butyl-4-methylphenol.
22. The formulation of any one of clauses 1 to 21, which comprises a complex comprising a transition metal ion and the chelant.
23. The formulation of clause 22, wherein the complex comprises a manganese, iron, cobalt, vanadium or copper ion.
24. The formulation of clause 22, wherein the complex comprises a manganese or iron ion.
25. The formulation of any one of clauses 1 to 21, which formulation comprises less than 0.001 % by weight of ions of each of manganese, iron, cobalt, vanadium and copper.
26. A method of preparing a formulation as defined in any one of clauses 1 to 25, the method comprising contacting a composition comprising an alkyd-based resin with a composition comprising a chelant of formula (I).
27. The method of clause 26 wherein the formulation is as defined in clause 25.
28. The method of clause 27 further comprising contacting the formulation with a source of transition metal ions.
29. The method of clause 28, wherein the transition metal ions are manganese, iron, cobalt, vanadium or copper ions.

28

30. The method of clause 28, wherein the transition metal ions are manganese or iron ions.

31. The method of clause 29 or clause 30 wherein a solution of transition metal ions
5 is contacted with the formulation.

32. The method of any one of clauses 28 to 31, wherein the formulation is contacted with an optionally hydrated salt selected from the group consisting of MnCl_2 , FeCl_2 , FeCl_3 , MnBr_2 , $\text{Mn}(\text{NO}_3)_2$, $\text{Fe}(\text{NO}_3)_3$, MnSO_4 , FeSO_4 , $(\text{Fe})_2(\text{SO}_4)_3$,
10 $\text{Mn}(\text{acetylacetonate})_2$, $\text{Fe}(\text{acetylacetonate})_2$, $\text{Mn}(\text{acetylacetonate})_3$, $\text{Fe}(\text{acetylacetonate})_3$, $\text{Mn}(\text{R}_5\text{COO})_3$, $\text{Fe}(\text{R}_5\text{COO})_3$, $\text{Mn}(\text{R}_5\text{COO})_2$ and $\text{Fe}(\text{R}_5\text{COO})_2$, wherein each R_5 is independently a C_1 - C_{24} alkyl.

33. The method of clause 32, wherein the formulation is contacted with an
15 optionally hydrated salt selected from the group consisting of MnSO_4 , MnCl_2 , MnBr_2 , $\text{Mn}(\text{NO}_3)_2$, $\text{Mn}(\text{CH}_3\text{COO})_3$, $\text{Mn}(\text{CH}_3\text{COO})_2$, $\text{Mn}(\text{acetylacetonate})_3$, $\text{Mn}(\text{acetylacetonate})_2$, $\text{Mn}(\text{2-ethylhexanoate})_2$, $\text{Mn}(\text{naphthenate})_2$, $\text{Mn}(\text{neodecanoate})_2$, $\text{Mn}(\text{octanoate})_2$, FeCl_2 , FeCl_3 , FeSO_4 , $\text{Fe}(\text{acetylacetonate})_3$, $\text{Fe}(\text{NO}_3)_3$, $\text{Fe}(\text{2-ethylhexanoate})_2$, $\text{Fe}(\text{neodecanoate})_2$, $\text{Fe}(\text{2-ethylhexanoate})_3$, $\text{Fe}(\text{naphthenate})_2$ and $\text{Fe}(\text{neodecanoate})_3$.

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34. The method of clause 26, wherein the composition comprising the chelant of formula (I) comprises a complex as defined in any one of clauses 22 to 24.

35. A composition resultant from curing of a formulation as defined in any one of
25 clauses 22 to 24.

36. A chelant of formula (I), as defined in clause 1, which is 6-{N,N-bis(pyridin-2-ylmethyl)amino}-1,4,6-trimethyl-1,4-diazacycloheptane or 6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane.

30

37. A kit comprising a formulation as defined in clause 25 and, separately, a composition comprising transition metal ions selected from the group consisting of manganese, iron, cobalt, vanadium and copper ions.

35 38. The kit of clause 37, wherein the transition metal ions are manganese ions or iron ions.

39. The kit of clause 38, wherein the ions are provided as a salt as defined in clause 32 or clause 33.

40. A method comprising applying to a substrate a formulation as defined in any one of clauses 22 to 24.

The non-limiting examples below more fully illustrate the embodiments of this invention.

10 **EXPERIMENTAL**

6-amino-1,4,6-trimethyl-1,4-diazacycloheptane, 6-dimethylamino-1,4,6-trimethyl-1,4-diazacycloheptane, 6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4,6-trimethyl-1,4-diazacycloheptane, 6-amino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane, and 6-dimethylamino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane have been prepared as disclosed in WO 01/85717 A1. 6-{N,N-bis(pyridin-2-ylmethyl)amino}-1,4,6-trimethyl-1,4-diazacycloheptane and 6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane may be prepared analogously.

20 All ligands used in the tests have been abbreviated as follows:

L1: 6-dimethylamino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;

L2: 6-amino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;

L3: 1,4,6-trimethyl-6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-diazacycloheptane;

L4: 6-dimethylamino-1,4,6-trimethyl-1,4-diazacycloheptane.

25 L5: 6-amino-1,4,6-trimethyl-1,4-diazacycloheptane; and

L6: 1,4,6-trimethyl-6-(pyridin-2-ylmethyl)amino)-1,4-diazacycloheptane.

The alkyd resin (catalogue number A/1552/15; an alkyd resin solution of 70 wt% in white spirits) was obtained from Acros Organics. Manganese(II) acetate tetrahydrate and manganese(III) acetate dihydrate were obtained from Aldrich. A commercial batch of manganese (II) 2-ethylhexanoate, 40% solution in mineral spirits, 6% manganese, ex Strem Chemicals Inc. has been used. This sample will be abbreviated as Mn(EH)₂ in the section below. A commercial batch of iron naphthenate, 40% in mineral spirits, 6% iron, ex Strem Chemicals Inc. has been used. This sample will be abbreviated as Fe(napht)₂ in the section below.

L1/Mn(EH)₂

15.45 mg of L1 dissolved in 0.05 mL ethanol was mixed with 41.7 mg Mn(2-EH)_2 and left in a closed vial for 15 minutes at room temperature (a 1:1 molar ratio Mn:L1 was present in the solution). Then 5 g of alkyd resin was added and the mixture was stirred manually for ca. 1 minute. This led to a Mn level of 0.05 wt-% with respect to resin
 5 solution (or 0.07 wt-% with respect to solid resin). The next day, 150 mg of this mixture was equally spread out on 10 cm² of a glass plate. Dryness of the film was determined every 30 minutes by running a needle through the film. 'Dry' was defined when the needle could not run through the film anymore, but gave wrinkling of the surface. A drying time of 4 h was measured.

10

L1/Fe(napht)₂

The experiment as described above was repeated with 15.20 mg of L1 dissolved in 0.05 mL ethanol and the Mn(2-EH) was replaced by 41.7 mg Fe(napht)_2 . This led to a drying time of 6.5 h.

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L2/Mn(EH)₂

The experiment as described above was repeated, except for using 14.17 mg of L2 dissolved in 0.05 mL ethanol. A drying time of 4.4 h was measured.

20

L2/Fe(napht)₂

The experiment as described for the Mn/L2 mixture was repeated, except 13.94 mg L2 was dissolved in 0.05 mL ethanol and the Mn(2-EH) was replaced by 41.7 mg Fe(napht)_2 . A drying time of 3.9 h was measured.

25

L3/Mn(EH)₂

The experiment as described for L1/Mn(EH)₂ above was repeated, except for using 11.94 mg of L3 dissolved in 0.05 mL ethanol. A drying time of 2 h was measured.

L3/Fe(napht)₂

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The experiment as described for the Mn/L3 mixture was repeated, except 11.75 mg L3 was dissolved in 0.05 mL ethanol and the Mn(2-EH) was replaced by 41.7 mg Fe(napht)_2 dissolved in 0.05 mL ethanol. A drying time of 2.5 h was measured.

L4/Mn(EH)₂

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The experiment as described for L1/Mn(EH)₂ above was repeated, except for using 8.43 mg of L4 dissolved in 0.05 mL ethanol. A drying time of 3.8 h was measured.

L4/Mn(CH₃COO)₂·4H₂O

The experiment as described for L4/Mn(EH)₂ above was repeated, except for using 11.15 mg of Mn(CH₃COO)₂·4H₂O dissolved in 0.10 mL ethanol. A drying time of 5 h was measured.

5

L4/Mn(CH₃COO)₃·2H₂O

The experiment as described for L4/Mn(EH)₂ above was repeated, except for using 12.20 mg of Mn(CH₃COO)₃·2H₂O dissolved in 0.20 mL ethanol. A drying time of 4 h was measured.

10

L5/Mn(EH)₂

The experiment as described for L1/Mn(EH)₂ above was repeated, except for using 7.16 mg of L5 dissolved in 0.05 mL ethanol. A drying time of 5.1 h was measured.

15 **L6/Mn(EH)₂**

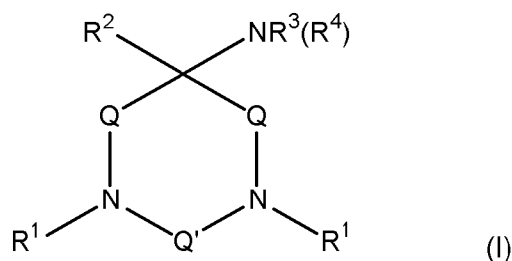
The experiment as described for L1/Mn(EH)₂ above was repeated, except for using 11.3 mg of L6 dissolved in 0.05 mL ethanol. A drying time of 6.4 h was measured.

Comparative experiments

- 20 Experiments L1/ Mn(EH)₂ and L1/ Fe(napht)₂ were repeated as described above, except now no L1 was added (so only Mn(EH)₂ and Fe(napht)₂ were tested on their drying behaviour. No drying was observed within 8 h.

The embodiments of the invention in which an exclusive property or privilege is claimed are defined as follows:

1. A formulation comprising an oxidatively curable alkyd-based resin and a complex, said complex comprising a transition metal manganese or iron ion and a chelant of the formula (I):



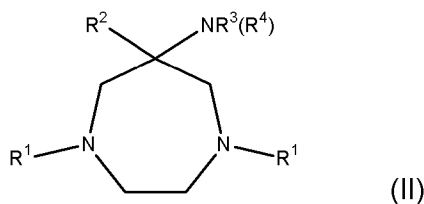
wherein:

each of -R¹, -R², -R³ and -R⁴ independently represents -H, -C₁-₂₄alkyl, -C₆-₁₀aryl or a group comprising a heteroatom capable of coordinating to a metal ion;

Q represents methylene or ethylene, in which one or more hydrogen atoms are optionally independently replaced with a C₁-₂₄ alkyl group or a C₆-₁₀ aryl; and

Q' represents ethylene or n-propylene, in which one or more hydrogen atoms are optionally independently replaced with a C₁-₂₄ alkyl group or a C₆-₁₀ aryl.

2. The formulation of claim 1, wherein the chelant is of formula (II):



wherein:

each -R¹ is independently -H, -C₁-₂₄alkyl, -C₆-₁₀aryl or pyridine-2ylmethyl, wherein aryl and pyridinyl are independently optionally substituted with C₁-₄alkyl;

-R² represents -H or -CH₃; and

each -R³ and -R⁴ is independently -C₁-₂₄alkyl, -C₆-₁₀aryl or pyridine-2ylmethyl, wherein aryl and pyridinyl are independently optionally substituted with C₁-₄alkyl.

3. The formulation of claim 1, wherein the chelant is:
 6-dimethylamino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;
 6-amino-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane;
 1,4,6-trimethyl-6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-diazacycloheptane;
 6-amino-1,4,6-trimethyl-1,4-diazacycloheptane;
 6-dimethylamino-1,4,6-trimethyl-1,4-diazacycloheptane;
 1,4,6-trimethyl-6-(pyridin-2-ylmethyl)amino)-1,4-diazacycloheptane;
 6-{N,N-bis(pyridin-2-ylmethyl)amino}-1,4,6-trimethyl-1,4-diazacycloheptane; or
 6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane.
4. The formulation of claim 3, wherein the chelant is 6-{N,N-bis(pyridin-2-ylmethyl)amino}-1,4,6-trimethyl-1,4-diazacycloheptane or 6-{N-(pyridin-2-ylmethyl)-N-methylamino}-1,4-bis(pyridine-2-ylmethyl)-6-methyl-1,4-diazacycloheptane.
5. The formulation of any one of claims 1 to 4, which formulation comprises less than 0.001 % by weight of ions of each of manganese or iron.
6. A method of preparing a formulation as defined in any one of claims 1 to 5, the method comprising contacting a composition comprising the oxidatively curable alkyd-based resin with a composition comprising the chelant of formula (I), the method further comprising contacting the formulation with a source of manganese or iron ions.
7. The method of claim 6, wherein the source of manganese or iron ions is a hydrated salt comprising MnCl_2 , FeCl_2 , FeCl_3 , MnBr_2 , $\text{Mn}(\text{NO}_3)_2$, $\text{Fe}(\text{NO}_3)_3$, MnSO_4 , FeSO_4 , $(\text{Fe})_2(\text{SO}_4)_3$, $\text{Mn}(\text{acetylacetonate})_2$, $\text{Fe}(\text{acetylacetonate})_2$, $\text{Mn}(\text{acetylacetonate})_3$, $\text{Fe}(\text{acetylacetonate})_3$, $\text{Mn}(\text{R}_5\text{COO})_3$, $\text{Fe}(\text{R}_5\text{COO})_3$, $\text{Mn}(\text{R}_5\text{COO})_2$ or $\text{Fe}(\text{R}_5\text{COO})_2$, wherein each R_5 is independently a $\text{C}_1\text{-C}_{24}$ alkyl.
8. A composition resultant from curing of a formulation as defined in any one of claims 1 to 5.

9. A kit comprising a formulation as defined in any one of claims 1 to 5 and a composition comprising the transition metal manganese or iron ions.
10. A method comprising applying to a substrate a formulation as defined in any one of claims 1 to 5.