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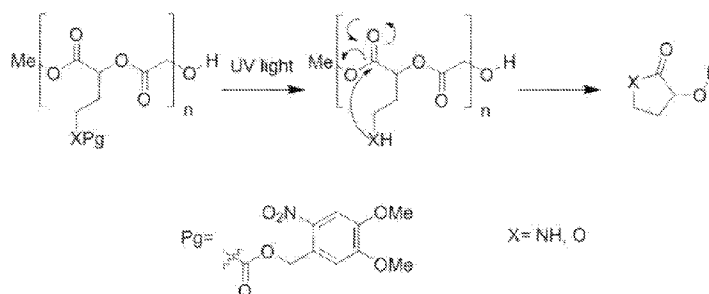
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(54) Title: STIMULUS-RESPONSIVE POLY(LACTIC-CO-GLYCOLIC)-BASED POLYMERS AND NANOPARTICLES FORMED THEREFROM



SCHEME 1

FIG. 1A

(57) Abstract: PLGA-based polymers include pendant nucleophiles protected with photocleavable protecting groups. Upon deprotection, the polymers degrade rapidly via intramolecular cyclization into small molecules. The polymer may be formulated as a nanoparticle, with an encapsulated payload, which may be an imaging agent, a bioactive agent or a pharmaceutical agent.

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**STIMULUS-RESPONSIVE POLY(LACTIC-CO-GLYCOLIC)-BASED
POLYMERS AND NANOPARTICLES FORMED THEREFROM****RELATED APPLICATIONS**

This application claims the benefit of the priority of U.S. Provisional
5 Application No. 62/144,845, filed April 8, 2015, which is incorporated herein by
reference in its entirety.

GOVERNMENT RIGHTS

This invention was made with government support under Grant No.
1DP2OD006499-01 awarded by the National Institutes of Health. The government
10 has certain rights in the invention.

FIELD OF THE INVENTION

The present invention relates polymers that rapidly degrade on demand, and
more particularly to polymers incorporating sidechains that allow cyclization after
removal of stimulus-responsive protecting groups.

15 BACKGROUND OF THE INVENTION

Stable polymers that depolymerize rapidly on demand upon application of a
specific stimulus are of great interest for a variety of industrial applications, such as
patterning, cosmetics, agriculture and electronics, and in biomedical applications such
as tissue engineering, tissue adhesives, and drug delivery. Despite this interest, few
20 synthetic polymers have been identified that have the ability to degrade with high
sensitivity in response to specific stimuli. Most current degradable materials are
unresponsive to the often subtle changes found in biological systems or, in the case of
photodegradable polymers, require long, intense irradiation that may not be
biologically compatible. The ongoing need for specifically engineered polymers is
25 clearly seen in the broad use of unresponsive poly(lactic-co-glycolic acid) (PLGA) in
current medical materials.

The emerging technology of nanoparticles packaging offers a way to package
and deliver compounds of interest that offers a number of advantages. Nanoparticles
can be synthesized and/or assembled so as to enclose other compounds of interest.
30 Thus, nanoparticles can serve to protect compounds of interest by sequestration and/or

encapsulation. Nanoparticulate media involved in this approach include nano- and microgels, nano- and microspheres, polymer micelles, and polymerized liposomes. Retention of the active compound in the nanocarriers is achieved by physical entrapment or by thermodynamic forces such as hydrophobic interactions.

5 Non-limiting examples of compounds of interest for delivery via nanoparticles to an area of interest, such as tumor tissue, include bioactive agents, pharmaceutical agents, or imaging agents. Nanoparticles may be then signaled to release their contents via externally-applied signals and/or signals present at the area of interest. In some examples, nanoparticles may be delivered systematically to a patient, while releasing
10 of the contents of the nanoparticle at a focused area of interest within a living organism. Non-limiting examples of the contents of nanoparticles, i.e., payloads, include pharmaceutical agents, drugs, antibodies, and/or labeling compounds.

Nanoparticle packaging can also improve the effectiveness of bioactive agents and/or pharmaceutical agents. In some nanoparticle designs, the serum stability of
15 bioactive agents and/or pharmaceutical agents can be enhanced and solubility limitations bypassed. Thus, nanoparticle packaging circumvents the vulnerability of bioactive agents and/or pharmaceutical agents, for example, to a reduction in efficacy due to bioavailability problems, e.g., solubility and/or stability. Moreover, such carriers can also serve to minimize undesirable side effects by reducing systemic
20 exposure to drugs and/or by decreasing their necessary dosage. In addition, encapsulating bioactive agents and/or imaging agents may protect them from sequestration and/or renal clearance.

Nanoparticles also offer the potential, at least, for targeted delivery of their payloads to specific areas of interest within a patient. For example, an affinity reagent
25 attached externally to nanoparticles enables an increase the concentration of such nanoparticles at their intended location. An example of such an affinity reagent is an antibody. Modifying the nanoparticles, and not the payload itself, avoids direct modification of the enclosed bioactive agent while improving its targeting and therefore obviates concerns about changing the activity of the bioactive agent.

30 Nanoparticles may be designed to be capable of a controlled and rapid triggered response to physiological events and/or conditions. Such physiological events and/or conditions may include changes in extracellular pH, temperature and

reactive oxygen species. Therefore, nanoparticles capable of such a triggered response may be useful in the delivery of therapeutics and diagnostics to diseased cells and tissue. (See, e.g., Farokhzad, et al., (2006) *Expert Opinion on Drug Delivery*, 3, 311-324; Farokhzad & Langer (2009) *ACS Nano*, 3, 16-20; Ferrari, (2005) *Nat. Rev. Cancer* 5, 161-171; Ganta, et al., (2008) *J. Control. Release*, 126, 187-204; Langer (1990) *Science*, 249, 1527-33; LaVan, et al. (2003) *Nat. Biotechnol.* 21, 1184-1191; Whitesides (2003) *Nat. Biotechnol.* 21, 1161-1165; and Zhang et al. (2008) *Clinical Pharmacology and Therapeutics*, 83, 761-9).

Additionally, encapsulation within nanoparticles constructed from biodegradable polymers can allow bioactive agents to be delivered to the cytosol of diseased cells via endosomes and cytosolic release (Lewis (1990) *Drugs and the Pharmaceutical Sciences*, Vol. 45: *Biodegradable Polymers as Drug Delivery Systems*, Chasin & Langer, Eds.; Marcel Dekker, pp 1-42; Panyam & Labhasetwar (2003) *Adv. Drug Delivery. Rev.*, 55, 329-347; and Shenoy, et al. (2005) *Pharm. Res.*, 22, 2107-14.). Cytosolic delivery is particularly challenging and can be a major hurdle for effective therapeutic delivery (Vasir & Labhasetwar (2007) *Adv. Drug Delivery. Rev.*, 59, 718-728; and Mescalchin et al. (2007) *Expert Opin. Biol. Ther.*, 7, 1531-1538). Burst-degrading drug delivery systems hold promise in achieving increased cytosolic release through elevated osmotic pressure within the endosomes (Sonawane, et al. (2003) *J. Biol. Chem.* 2003, 278, 44826-31; and Hu, et al. (2007) *Nano Lett.* 7, 3056-64).

In the past, nanoparticles have been developed from hydrogels utilizing ketal crosslinks. However the payloads of such nanoparticles are usually limited to large water-soluble macromolecules. Unfortunately, with nanoparticles such as these, significant unwanted degradation occurs at physiological pH values over time (Cohen, et al. (2008) *Bioconjug. Chem.*, 19, 876-81). Similarly, hydrophobic polyketals can encapsulate both hydrophobic and hydrophilic payloads, however, as nanoparticles they no longer undergo rapid acid catalyzed hydrolysis unless fully hydrated (Yang, et al. (2008) *Bioconjug. Chem.*, 19, 1164-1169).

Formulation of nanoparticles from polymers may provide them with a hydrophobic character. However, this dramatically slows down their hydrolysis degradation kinetics as degradation only occurs slowly by a surface erosion

mechanism (Heffernan, et al. (2009) *Biomaterials*, 30, 910-918; Heffernan & Murthy (2005) *Bioconjug. Chem.*, 16, 1340-1342; Paramonov, et al. (2008) *Bioconjug. Chem.*, 19, 911-919).

There is growing interest in polymeric biomaterials that can be remotely
5 disassembled in a controlled fashion with an external stimulus, but are otherwise stable
under physiological conditions (Wang, W.; Alexander, C. *Angew. Chem. Int. Ed.*,
2008, 47, 7804-7806). Various internal and external stimuli, such as specific enzymes,
temperature, ultrasound, and electromagnetic radiation, have been exploited as release
mechanisms. (See, e.g., Veronese, et al. (2005) *Bioconjugate Chem.* 16, 775-784;
10 Chung, et al. (1999) *Controlled Release*, 62 (1-2), 115-27; Liu, et al. (2005)
Biomaterials, 26, 5064-5074; Na, et al. (2006) *Eur. J. Pharm. Sci.*, 27, 115-122; Gao,
et al. (2005) *Controlled Release*, 102, 203-22; Nelson, et al. (2002) *Cancer Research*,
62, 7280-83); and PCT Publication WO 2011/038117 A2, Almutairi et al.)

Nanoparticles composed of synthetic polymers such as poly(lactic-co-glycolic
15 acid) (PLGA) are safe and attractive methods for DNA delivery applications and have
been used in several studies. PLGA polyesters can be degraded by hydrolysis,
facilitating their widespread use in medicine and biomedical research. Their
dependence on slow hydrolysis makes for long degradation times (half-life of one year
in vivo), thus limiting their applicability. While degradation can be sped up by
20 copolymerization with more hydrophilic monomers; degradation is still too slow for
triggered release or degradation.

Poly(lactide) (PLA) and poly(D,L-lactide-co-glycolide) (PLGA) have been
thoroughly investigated as drug delivery vehicles because of their slow hydrolytic
degradation to largely biologically innocuous substances, but these polymers offer
25 minimal control over degradation. Molecular engineering of the PLGA structures
would accelerate degradation rates and allow triggered degradation.

SUMMARY OF THE INVENTION

In an exemplary embodiment, a series of polymers incorporating sidechains
allow cyclization after removal of stimulus-responsive protecting groups such that
30 exposure to the stimulus (e.g., light) would expose a pendant nucleophile to trigger
intramolecular cyclization. This causes breaks in the polymer backbone and

degradation into cyclic small molecules. The backbone itself will also degrade slowly via hydrolysis so that even with incomplete triggered degradation, the polymers still break down in an aqueous environment. These polymers add triggered degradation to the array of properties now found in the growing field of functional hydrolytically
5 degradable PLGA-type polymers.

In one aspect of the invention, a novel polymeric structure is provided to enable rapid on-demand degradation and expand the library of polymers that degrade by cyclization. We have synthesized, by ring-opening polymerization (ROP), a series of PLGA-based polymers containing pendant nucleophiles protected with
10 photocleavable protecting groups. Upon deprotection, of the polymers degrade rapidly via intramolecular cyclization into small molecules. Nanoparticles formulated from these polymers undergo rapid structural changes in response to UV light.

The inventive approach has been demonstrated with polymers bearing ortho-nitrobenzyl groups, yielding light-degradable polymers. ¹H NMR spectra of model
15 polymers of the same backbone were used to demonstrate degradation into cyclic small molecules, and fluorescence quenching of nanoparticle-encapsulated (by single emulsion) Nile red and dynamic light scattering confirmed light-triggered release and degradation of particles.

In one aspect of the invention, a composition comprises a polymer having a
20 poly(lactide-*co*-glycolide)-type backbone and pendant nucleophiles protected by a stimulus-responsive protecting group, wherein the protecting group is configured to deprotect upon exposure to a stimulus to facilitate degradation by intramolecular cyclization. The protecting group may be photocleavable, where the stimulus may be UV light. In some embodiments, the protecting group is an ortho-nitrobenzyl (ONB)
25 protecting group. The polymer may be formulated as a nanoparticle, and further comprise a payload encapsulated within the nanoparticle. The payload may be an imaging agent, a bioactive agent or a pharmaceutical agent. The pendant nucleophile may be an amine, an alcohol or a thiol.

In another aspect of the invention, a composition comprises a polymer
30 comprising a polymer backbone comprising a poly(lactide-*co*-glycolide); and pendant nucleophiles linked to the backbone by photolabile protecting group, wherein the protecting group is configured to deprotect upon exposure to irradiation to facilitate

degradation of the nucleophiles to five membered rings by intramolecular cyclization. In some embodiments, the irradiation is UV light. The protecting group is preferably an ortho-nitrobenzyl (ONB) protecting group. The polymer may be formulated as a nanoparticle, and further comprise a payload encapsulated within the nanoparticle.

5 The payload may be an imaging agent, a bioactive agent or a pharmaceutical agent. The pendant nucleophile may be an amine, an alcohol or a thiol.

The inventive polymers have applications in on-demand and controlled drug delivery where the favorable properties of hydrophobic polyesters (processibility and fabrication into particles, fibers, implants, etc.) can be combined with the favorable

10 properties of hydrophilic polyesters (good biodegradation). Moreover, as the versatile design of this system allows easy replacement of the triggering group, the presented strategy may have a broad impact on polymer science.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1A shows the polymer degradation mechanism according to an

15 embodiment of the invention; FIG. 1B is a diagram of the light-degradation process according to an embodiment of the invention.

FIG. 2A shows the synthesis of three monomers (**3**, **8** and **14**) according to one embodiment of the invention (Scheme 2); FIGs. 2B-2D show synthesis of monomers **3**, **8** and **14**, respectively, with added detail.

20 FIG. 3 shows preparation of three exemplary light-degradable polymers (polymers **16**, **17** and **18**) according to another embodiment of the invention (Scheme 3).

FIGs. 4A-4C are UV absorbance spectra upon irradiation with UV light of polymer **16**, polymer **17**, and polymer **18**, respectively, with a spectrum recorded at 0

25 minutes and each minute thereafter for 18 minutes total. FIG. 4D is a plot showing the change in absorbance at 346 nm upon irradiation for the three polymers.

FIGs. 5A and 5B show GPC traces of polymer **17** and polymer **18**, respectively, following 15 min irradiation (1 mW/cm²) and subsequent incubation for the specified times at 37 °C. The irradiation was relatively brief for this concentration

30 so the cyclization chemistry could be readily observed.

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FIGs. 6A-C are GPC traces for polymer **16**, polymer **17** and polymer **18**, respectively, without irradiation.

FIGs. 7A-C are GPC traces for polymer **16**, polymer **17** and polymer **18**, respectively, following 15 min. irradiation.

5 FIG. 8 shows the synthesis of the model polymer in an embodiment of the invention (Scheme 4).

FIG. 9A shows deprotection and subsequent degradation of model polymer **22** to likely cyclic degradation products and ^1H NMR spectra of degradation products of polymer **22** (upper) and compound **27** (lower); FIG. 9B shows deprotection and
10 subsequent degradation of model polymer **26** to likely cyclic degradation products and ^1H NMR spectra of degradation products of polymer **26** (upper) and compound **28** (lower).

FIGs. 10A-10C are TEM micrographs of NP **16**, NP **17**, and NP **18**, respectively. Scale bar = 100 nm.

15 FIG. 11A is a TEM micrograph of NP **16** after irradiation at 320-390 nm (0.181 W/cm^2) for 15 sec and overnight incubation at 37°C (scale bar = 100 nm); FIG. 11B is a plot of the DLS count rate of NP **16**, NP **17** and NP **18** irradiated at 320-390 nm (0.181 W/cm^2) for 15 s and incubated at room temperature in 1X PBS (pH 7.4); FIG. 11C plots the results for non-irradiated NP **16**, NP **17** and NP **18** incubated at
20 room temperature in 1X PBS (pH 7.4); and FIG. 11D plots the DLS count rate for PLGA particles irradiated with the above conditions and not irradiated incubated at room temperature in 1X PBS (pH 7.4).

FIGs. 12A and 12B are TEM micrographs of NP **17** following irradiation; FIGs. 12C and 12D show NP **18** after irradiation.

25 FIGs. 13A-C are SEM images of nanoparticles NP **16**, NP **17** and NP **18**, respectively, each encapsulating Nile red.

FIGs. 14A and 14B are plots showing fluorescence quenching of Nile red encapsulated polymeric nanoparticles, irradiated with UV light (1 mW/cm^2) and non-irradiated with incubation at 37°C , respectively.

DETAILED DESCRIPTION OF EXEMPLARY EMBODIMENTS

According to embodiments of the invention, polymers incorporating sidechains allow cyclization after removal of stimulus-responsive protecting groups such that exposure to the stimulus (e.g., light) exposes a pendant nucleophile to trigger intramolecular cyclization. This causes breaks in the polymer backbone and degradation into cyclic small molecules. The backbone itself will also degrade slowly via hydrolysis so that even with incomplete triggered degradation, the polymers still break down in an aqueous environment. These polymers add triggered degradation to the array of properties now found in the growing field of functional hydrolytically degradable PLGA-type polymers.

Table 1 below provides a listing of abbreviations/acronyms and definitions used throughout the written description:

<u>Abbreviation</u>	<u>Definition</u>
AcOH	Acetic Acid
DLS	Dynamic light scattering
DMAP	4-Dimethylaminopyridine
DMF	Dimethylformamide
GPC	Gel permeation chromatography
HRMS	High-resolution mass spectrometry
MeCN	Acetonitrile
NMR	Nuclear magnetic resonance
NP	Nanoparticle
ONB	<i>ortho</i> -nitrobenzyl
PBS	Phosphate Buffered Saline
PDI	Polydispersity Index
PLGA	Poly(lactic- <i>co</i> -glycolic acid)
PMMA	Poly(methyl methacrylate)
ppm	parts per million
ROP	ring-opening polymerization
TBAF	Tetra- <i>n</i> -butylammonium Fluoride
TBDPSCI	tert-Butyldiphenylchlorosilane
TEM	Transmission electron microscopy
TFA	Trifluoroacetic Acid
THF	tetrahydrofuran
UV	Ultraviolet

Table 1

In embodiments of the invention, polymers were designed to include (Scheme 1). An *ortho*-nitrobenzyl (ONB) protecting group, which degrades in response to UV light, was selected as the photocleavable moiety because this variety of protecting group is well-studied, commonly used in similar applications, and readily available.

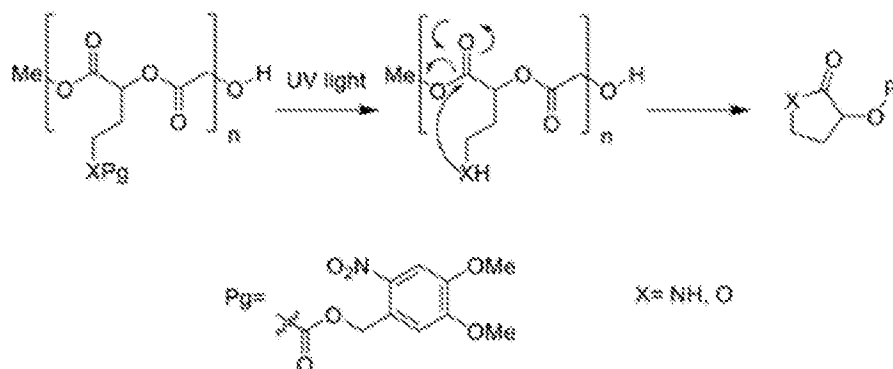
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UV-degradable polymeric particles and other materials have been employed for biologically relevant purposes. These polymers add triggered degradation to the array of properties now found in the growing field of functional hydrolytically degradable poly(α -hydroxyl acid)s. The synthesis of the described polymers by a controlled ROP
 5 allows study of these novel backbones and shows their potential for direct use in degradable biomedical devices such as polymeric nanoparticles for drug delivery.

Polymeric nanoparticles composed of these polymers, when triggered, should rapidly degrade and release encapsulated molecules. Upon irradiation the component polymers immediately become more hydrophilic, allowing water into the particles, facilitating hydrolysis of the backbone, and would degrade rapidly by intramolecular
 10 cyclization. FIG. 1B diagrammatically illustrates the process for degradation of nanoparticles formulated from the polymers, where UV light triggers degradation into 5-membered rings and photocages encapsulating a payload, including, but not limited to, imaging agents, bioactive agents or pharmaceutical agents. The PLGA-based
 15 backbone also ensures complete degradation into small molecules through slower ester hydrolysis, facilitating their clearance, even without complete removal of pendant photocages.

Scheme 1 (also provided as FIG. 1A) shows the polymer degradation mechanism, where removal of the photolabile protecting group frees the pendant
 20 nucleophile to cleave the backbone ester and form a five membered ring.

SCHEME 1



All chemicals and solvents were purchased from Sigma-Aldrich (St. Louis, MO) and used as received unless otherwise specified. Compound **19**, 1-[3,5-
 25 bis(trifluoromethyl)phenyl]-3-[(1R,2R)-(-)-2-(dimethylamino)cyclohexyl]thiourea

(R,R-TUC), was purchased from Strem Chemicals (Newburyport, MA) and used as received. Anhydrous solvents were acquired from a solvent purification system (LC Technology Solutions, Inc. (Salisbury, MA), SP-1). Silica gel flash column chromatography was performed using an automated CombiFlash[®] Rf200 system.

5 Polymer molecular weights and degradation were determined by gel permeation chromatography using a Waters e2695 instrument with a series of Styragel HR4 and Styragel HR2 columns in DMF with 0.01% LiBr at 37 °C. Monodisperse PMMA standards were used to determine the molecular weight and PDI of polymers. ¹H NMR and ¹³C NMR spectra were obtained using a Varian spectrometer working at 600

10 MHz and 150 MHz, respectively. Chemical shifts (δ) are reported in ppm relative to TMS, and coupling constants (J) are reported in Hertz. High-resolution mass spectra were acquired using an Agilent 6230 ESI-TOFMS in positive ion mode. Irradiation with UV light was done with a Luzchem LZ-ORG photoreactor equipped with 8 UV-A lamps with a power of 1.35 mW/cm² and a 0.181 W/cm², $\lambda_{\text{ex}} = 320\text{-}90\text{ nm}$,

15 OmniCure S2000 Curing System. Particles were formulated using a Qsonica Sonicator 4000 and purified by tangential flow filtration using Millipore Pellicon XL, 500 kDa. Particles were characterized by DLS, Malvern Instruments Nanosizer, and scanning electron microscopy (TEM, Tenai Spirit). Fluorescence was measured using a Horiba Jobin FL-fluorimeter.

20 Three polymers with different pendant nucleophiles were synthesized: an amine, an alcohol, and a thiol. Preparation of a polymer series not only allows comparison of their degradation, but also examination of what chemistries are compatible with cyclization. Prior intramolecularly cyclizing polymers did not include alcohols as nucleophiles. Broadening the array of nucleophiles to include alcohols

25 expands the types of chemistry that can be used to trigger degradation. Degradation by cyclization does not occur with analogous thiol nucleophiles in this series. Polymer properties were also compared in polymeric particles, which could be used for drug delivery or other applications.

The synthetic routes for the three monomers **3**, **8** and **14** are shown in Scheme

30 2. The monomer with a pendant amine, monomer **3**, was readily prepared from **1**. The amine of **1** was first protected with the UV light-sensitive ortho-nitrobenzyl protecting group using triethylamine and compound **15**. Alcohol **2** was then acylated with

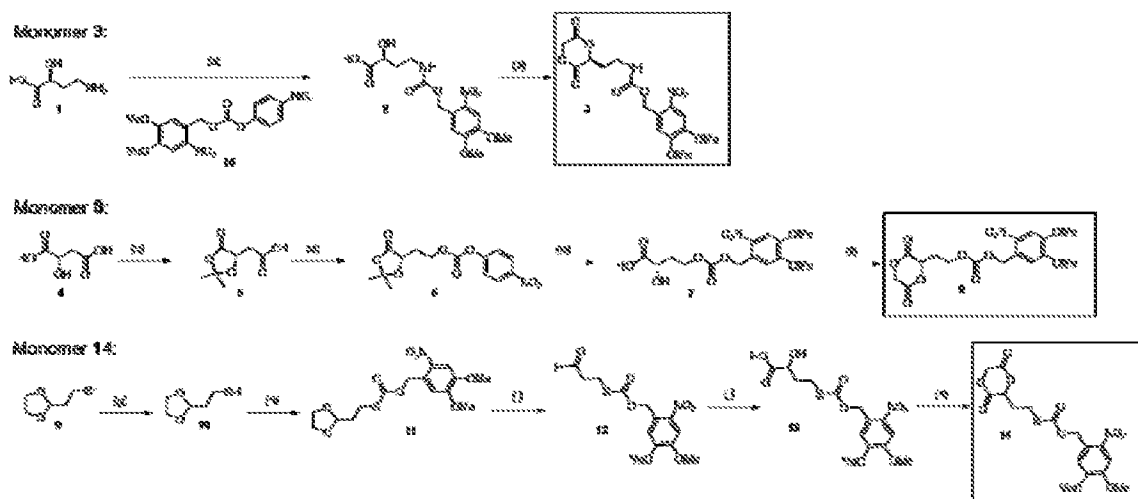
bromoacetyl bromide and subsequently cyclized with sodium bicarbonate to give dilactone **3** in a manner similar to that outlined by Pounder et al.

To obtain monomer **8** L-malic acid **4** was protected to form acetal **5** following an established procedure for D-malic acid. Acetal **5** was then reduced with borane to yield an alcohol that was immediately reacted with 4-nitrophenyl chloroformate to form carbonate **6**. Carbonate **6** was reacted with 4,5-dimethoxy-2-nitrobenzyl alcohol to install the light-sensitive protecting group. The acetal was then hydrolyzed to reveal the vicinal alcohol and acid of **7**. Compound **7** was cyclized to form the dilactone monomer **8** in a manner similar to that described for monomer **3**.

Synthesis of monomer **14** required a distinct route because the ketal protecting group proved too labile under the conditions required to generate a thiol from compound **5**. Instead, the protected thiol was formed by a more mild substitution reaction, a nucleophilic displacement of the bromine of compound **9** with sodium hydrosulfide, to yield thiol **10**. Thiol **10** was protected with the ortho-nitrobenzyl protecting group using compound **15** to yield compound **11**, the acetal of which was hydrolyzed to yield aldehyde **12**. Aldehyde **12** was treated with trimethyl silyl cyanide and zinc iodide to yield a cyanohydrin which was then hydrolyzed to afford acid **13**. Compound **13** was reacted with bromoacetyl bromide in the same manner as the previous two monomers to afford dilactone monomer **14**.

Scheme 2 (also shown as FIG. 2A) shows the synthesis of monomers with reagents and conditions: a) **15**, Et₃N, MeCN, reflux, 61%; b) (i) bromoacetyl bromide, Et₃N, MeCN, 0 °C; (ii) NaHCO₃, DMF, 45% over two steps, c) p-toluenesulfonic acid, 2,2-dimethoxypropane, CH₂Cl₂, 95%; d) (i) borane, THF 0 °C; (ii) 4-nitrophenyl chloroformate, pyridine, CH₂Cl₂, 0 °C, 48% over two steps; e) (i) 4,5-dimethoxy-2-nitrobenzyl alcohol, DMAP, pyridine, CH₂Cl₂, 0 °C, 68%; (ii) THF, H₂O, AcOH, 100%; f) (i) bromoacetyl bromide, Et₃N, MeCN, 0 °C; (ii) DMF, NaHCO₃, 58% over two steps. g) Sodium hydrosulfide, DMF, 0 °C, 80%; h) **15**, pyridine, DMAP, CH₂Cl₂, 0 °C, 85%; (i) THF, 1M HCl, 98%; j) (i) Trimethylsilyl cyanide, ZnI₂, CH₂Cl₂, 0 °C, 98%; (ii) HCl (Concentrated), 1,4-Dioxane, 50 °C, 59%; k) (i) bromoacetyl bromide, Et₃N, MeCN 0 °C; (ii) DMF, NaHCO₃, 41% over two steps.

SCHEME 2



FIGs. 2B-2D provide additional details for synthesis of monomers **3**, **8**, and **14**.

Referring to FIG. 2B (monomer **3**), for compound **2**, compound **1** (1.70 g, 14.3 mmol) and compound **15** (4.153 g, 10.93 mmol) were suspended in MeCN (170 mL). Et₃N (6.1 mL, 43.8 mmol) was added dropwise to the suspension. The reaction mixture was heated at reflux for 18 h. The reaction mixture was concentrated and purified by silica column (1:1 MeOH/ EtOAc) to yield compound **2** as a yellow solid (2.38 g, 61%).

High-resolution mass spectrometry (HRMS) was used to determine composition of compound **2**: C₁₄H₁₉N₂O₉; measured mass 358.1087; theoretical mass: 358.1085.

NMR Results: ¹H NMR (600 MHz, CDCl₃) δ 7.67 (s, 1H), 6.99 (s, 1H), 5.76-5.66 (bs, 1H) 5.46 (s, 2H), 4.21-4.15 (m, 1H), 3.97 (s, 3H), 3.93 (s, 4H), 3.48-3.32 (m, 2H), 2.12-1.98 (m, 1H), 1.91-1.80 (m, 1H). ¹³C NMR (151 MHz, CDCl₃) δ 162.9, 156.5, 153.8, 148.1, 139.6, 128.5, 110.1, 108.1, 63.6, 56.6, 56.5, 38.2, 36.7, 34.1.

For compound **3a**, compound **2** (900 mg, 2.51 mmol) was dissolved in MeCN (18 mL) and was chilled to 0 °C. Et₃N (0.39 mL, 2.76 mmol) was added dropwise. A solution of bromoacetyl bromide (0.24 mL, 2.76 mmol) in MeCN (9 mL) was added dropwise to the chilled solution of compound **3**. The reaction was quenched after 1 hour of stirring by adding 1 M HCl. The reaction mixture was extracted three times with ethyl acetate the combined organic was washed with brine, dried over MgSO₄

and was concentrated. The resulting orange oil was used in the next step without further purification.

Compound **3**: Compound **3a** (2.51 mmol), used without purification, was dissolved in DMF (37.5 mL) and dripped into a suspension of NaHCO₃ (316 mg, 3.76 mmol) in DMF (75 mL) over a period of 28 hours. The reaction was stirred for a further 8 hours. The reaction mixture was filtered and concentrated. The resulting oil was purified by silica column (1:3 Hex/ EtOAc) to yield compound **3** as a yellow solid (450 mg, 45%) over two steps.

Using HRMS, composition was C₁₆H₁₈N₂O₁₀Na; measured mass 421.0856; theoretical mass: 421.0854.

NMR results: ¹H NMR (600 MHz, CDCl₃) δ 7.69 (s, 1H), 6.99 (s, 1H), 5.49 (d, *J* = 14.4 Hz, 1H), 5.47 (d, *J* = 14.4 Hz, 1H), 5.15-5.08 (m, 2H), 4.95 (d, *J* = 16.2 Hz, 1H), 4.91 (d, *J* = 16.2 Hz, 1H), 3.99 (s, 3H), 3.96 (d, *J* = 12.4 Hz, 3H), 3.55 - 3.41 (m, 2H), 2.42 (m, 1H), 2.20 (m, 1H). ¹³C NMR (151 MHz, CDCl₃) δ 165.8, 164.2, 156.3, 148.5, 140.2, 127.5, 111.0, 108.4, 100.1, 73.6, 65.7, 64.1, 56.7, 56.6, 36.7, 31.2.

Referring to FIG. 2C, compound **5** was synthesized from compound **4**, L-malic acid, via a previously-described method using D-malic acid.

For compound **6**, compound **5** (929 mg, 5.3 mmol) was dissolved in THF (6 mL) in base washed glassware. The solution was chilled to 0 °C and a 1 M solution of borane in THF (8 mL) was added dropwise over 30 min. The reaction was allowed to warm to room temperature gradually and allowed to react for 3.5 h. The reaction mixture was quenched with methanol at -78 °C and was concentrated. The reaction mixture was dissolved once more in methanol and concentrated, then dissolved in ethyl acetate and concentrated a final time. The reaction mixture was dissolved in CH₂Cl₂ (8 mL) without purification. The solution was chilled to 0 °C and pyridine (1.401 mL, 17.4 mmol) was dripped in. A solution of 4-nitrophenyl chloroformate (2.338 g, 11.6 mmol) in CH₂Cl₂ (12 mL) was slowly dripped into the reaction mixture. The mixture was allowed to warm to room temperature and react for 12 hours. The reaction was quenched with distilled water, organic collected and aqueous extracted 2X more with CH₂Cl₂. The combined organic was washed with brine, dried over MgSO₄, and concentrated. Purified by silica column (CH₂Cl₂) to yield compound **6**

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as colorless oil (0.900 g, 47.7 %) HRMS: composition: $C_{14}H_{15}NO_8Na$; measured mass 348.0689; theoretical mass: 348.0690.

NMR results for compound **6**: ^1H NMR (600 MHz, $CDCl_3$) δ 8.31- 8.26 (m, 2H), 7.41 - 7.37 (m, 2H), 4.56 (dd, $J = 7.0, 4.8$ Hz, 1H), 4.54-4.49 (m, 1H), 4.48-4.43 (m, 1H), 2.38-2.31 (m, 1H), 2.29 - 2.21 (m, 1H), 1.65 (s, 3H), 1.58 (s, 3H). ^{13}C NMR (151 MHz, $CDCl_3$) δ 172.6, 155.5, 152.3, 145.5, 125.4, 122.0, 111.2, 70.8, 64.7, 30.4, 27.1, 25.8.

Still referring to FIG. 2C, for compound **7a**, 4,5-dimethoxy-2-nitrobenzyl alcohol (1.096 g, 5.15 mmol) and DMAP (62.9 mg, 0.515 mmol) were dissolved in CH_2Cl_2 (11 mL) and chilled to 0 °C. Pyridine (0.83 mL, 10.3 mmol) was dripped into the solution. A solution of compound **6** (1.676 g, 5.15 mmol) in DCM (13 mL) was added to the reaction mixture. The reaction mixture was warmed to room temperature and allowed to react for 16 hours. The reaction mixture was quenched with saturated $NaHCO_3$ solution and was extracted 3X with CH_2Cl_2 , the combined organic was washed with brine, dried over $MgSO_4$, then concentrated. Purified by silica column (CH_2Cl_2) to yield compound **7a** as yellow solid (1.392 g, 67.6%).

Using HRMS, the composition was determined to be $C_{17}H_{21}NO_{10}Na$; measured mass was 422.1055; theoretical mass: 422.1058.

NMR results for compound **7a**: ^1H NMR (600 MHz, $CDCl_3$) δ 7.73 (s, 1H), 7.10 (s, 1H), 5.62 (d, $J=14.4$ Hz, 1H), 5.58 (d, $J=14.4$ Hz, 1H), 4.52 (dd, $J = 6.9, 4.8$ Hz, 1H), 4.45 - 4.34 (m, 2H), 4.02 (s, 3H), 3.96 (s, 3H), 2.32 - 2.24 (m, 1H), 2.23 - 2.14 (m, 1H), 1.62 (s, 3H), 1.56 (s, 3H). ^{13}C NMR (151 MHz, $CDCl_3$) δ 172.7, 154.4, 153.8, 148.3, 139.6, 126.8, 111.0, 109.9, 108.1, 70.8, 66.4, 63.6, 56.7, 56.4, 30.6, 27.1, 25.8.

For compound **7**, compound **7a** (1.827 g, 4.57 mmol) was dissolved in THF/ H_2O /AcOH 1:1:1 (19 mL), the mixture was stirred for 118 hrs. The reaction mixture was concentrated then dissolved in CH_2Cl_2 . The CH_2Cl_2 was washed one time with 1 M HCl. The aqueous was extracted twice with CH_2Cl_2 . The combined organic was dried over $MgSO_4$ and concentrated to yield compound **7** as a yellow oil (1.64 g, 100%). Using HRMS, composition was $C_{14}H_{17}NO_{10}Na$. Measured mass was 382.0744; theoretical mass: 382.0745.

NMR analysis results for compound **7**: ^1H NMR (600 MHz, CDCl_3) δ 7.70 (s, 1H), 7.03 (s, 1H), 5.56 (s, 2H), 4.44 - 4.33 (m, 3H), 3.97 (s, 3H), 3.95 (s, 3H), 2.30-2.20 (m, 1H), 2.12 - 2.03 (m, 1H). ^{13}C NMR (151 MHz, CDCl_3) δ 177.9, 154.9, 153.9, 148.5, 139.8, 126.5, 110.4, 108.3, 67.1, 66.7, 64.3, 56.7, 56.5, 33.0.

5 For compound **8**, compound **7** (1.487 g, 4.14 mmol) was dissolved in acetonitrile (28 mL) and chilled to 0 °C. Et_3N (0.635 mL, 4.55 mmol) was dripped in followed by the dropwise addition of a solution of bromoacetyl bromide (0.918 g, 4.55 mmol) in MeCN (14 mL). The reaction was quenched with 1 M HCl, and extracted 3X with ethyl acetate. The combined organic was dried over MgSO_4 and concentrated.
10 The resulting orange oil was used without further purification. The oil was dissolved in DMF (63 mL) and dripped into a suspension of NaHCO_3 (521 mg, 6.21 mmol) in DMF (126 mL) over a period of 28 hours. The reaction was stirred for a further 8 hours. The reaction mixture was concentrated, then suspended in ethyl acetate. The suspension was filtered and the filtrate was collected and concentrated. The resulting
15 oil was purified by silica column (1:1 Hex/ EtOAc) then recrystallized (CH_2Cl_2 /Ether) to yield compound **8** as a yellow solid (0.952 g, 57.7%) over two steps.

NMR analysis results for compound **8**: ^1H NMR (600 MHz, CDCl_3) δ 7.73 (s, 1H), 7.05 (s, 1H), 5.60 (d, J = 14.4 Hz, 1H), 5.57 (d, J = 14.4 Hz, 1H), 5.11 (dd, J = 7.8, 4.2 Hz, 1H), 5.00 (d, J =16.2Hz, 1H), 4.94 (d, J = 16.2 Hz, 1H), 4.49 - 4.42 (m, 2H), 4.01 (s, 3H), 3.97 (s, 3H), 2.61 - 2.53 (m, 1H), 2.40-2.34 (m, 1H). ^{13}C NMR (151 MHz, CDCl_3) δ 165.4, 163.8, 154.5, 153.8, 148.7, 126.1, 110.8, 108.4, 100.1, 72.1, 67.0, 66.7, 62.8, 56.8, 56.6, 30.21.

Referring now to FIG. 2D, synthesis of monomer **14** begins with compound **10**, sodium hydrosulfide (2.477 g, 44.2 mmol), which was dissolved in DMF and
25 chilled to 0 °C. Compound **9** (2.59 mL, 22.1 mmol) was dripped in and the reaction was allowed to warm to room temperature and stirred 20 h. The reaction was quenched by adding water. The reaction mixture was extracted 4X with hexanes, after which the combined organic was dried over MgSO_4 and concentrated to yield compound **10** (2.384 g, 80.4%) as a colorless oil.

30 NMR analysis results for compound **10**: ^1H NMR (600 MHz, CDCl_3) δ 4.96 (t, J = 4.5 Hz, 1H), 4.00 - 3.92 (m, 2H), 3.90 - 3.81 (m, 2H), 2.64 (q, J = 15.1, 7.8 Hz,

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2H), 1.99 (dt, $J = 10.2, 4.2$ Hz, 2H), 1.48 (t, $J = 7.8$ Hz, 1H). ^{13}C NMR (151 MHz, CDCl_3) δ 103.1, 65.1, 38.2, 19.2.

For compound **11**, a suspension of compound **10** (4.225 g, 11.2 mmol) and DMAP (0.136 g, 1.1 mmol) was prepared in CH_2Cl_2 (24 mL) and chilled to 0 °C. Pyridine (1.799 g, 22.4 mmol) was dripped into the suspension. Compound **15** (1.5 g, 11.17 mmol) was dissolved in CH_2Cl_2 (15 mL) and dripped into the suspension. The suspension was allowed to come to room temperature and react for 16 hours, by which time all material was in solution. The reaction was quenched by diluting with CH_2Cl_2 then washing with 1 M HCl, deionized water, saturated sodium carbonate twice, and finally one wash with brine. The organic layer was dried over MgSO_4 and concentrated. The resulting oil was purified by silica column ($\text{CH}_2\text{Cl}_2/\text{MeOH}$ 9:1) to yield compound **11** as a yellow solid (3.554 g, 85.2 %).

NMR analysis results for compound **11**: ^1H NMR (600 MHz, CDCl_3) δ 7.73 (s, 1H), 7.00 (s, 1H), 5.67 (s, 2H), 4.96 (t, $J = 4.3$ Hz, 1H), 4.00 - 3.94 (m, 8H), 3.90 - 3.83 (m, 2H), 3.04 - 2.99 (m, 2H), 2.08 - 2.01 (m, 2H). ^{13}C NMR (151 MHz, CDCl_3) δ 170.8, 153.9, 148.4, 139.7, 127.1, 109.9, 108.3, 102.9, 65.7, 65.6, 65.2, 56.6, 56.5, 34.0, 25.6.

For compound **12**, compound **11** (2 g, 5.3 mmol) was dissolved in a 1:1 mixture of THF and 1 M HCl (40 mL). The reaction mixture was stirred for 5 days. The reaction mix was diluted with deionized water and extracted 3X with CH_2Cl_2 . The combined organic was dried over MgSO_4 and concentrated to yield Compound **12** (1.702 g, 97.8%) as a yellow oil.

HRMS analysis identified the composition: $\text{C}_{13}\text{H}_{15}\text{NO}_7\text{SNa}$. Measured mass was 352.0460; theoretical mass: 352.0461.

NMR analysis results for compound **12**: ^1H NMR (600 MHz, CDCl_3) δ 9.78 (s, 1H), 7.73 (s, 1H), 6.98 (s, 1H), 5.67 (s, 2H), 3.99 (s, 3H), 3.96 (s, 3H), 3.14 (t, $J = 6.6$ Hz, 2H), 2.92 (t, $J = 6.6$ Hz, 2H). ^{13}C NMR (151 MHz, CDCl_3) δ 199.6, 170.6, 153.9, 148.5, 139.8, 126.7, 110.0, 108.4, 66.9, 56.7, 56.6, 44.1, 23.6.

For compound **13a**, compound **15** (1.302 g, 3.95 mmol) and zinc iodide were dissolved in CH_2Cl_2 and chilled to 0 °C. Trimethylsilyl cyanide (0.989 mL, 7.91 mmol) was dripped in and the reaction was allowed to proceed for 1 hour. The reaction was then diluted with CH_2Cl_2 and washed 3X with saturated NaHCO_3

solution. The organic layer was dried over MgSO_4 and concentrated to yield compound **13a** (1.656 g, 97.7 %) as a yellow oil.

NMR analysis results for compound **12**: ^1H NMR (600 MHz, CDCl_3) δ 7.73 (s, 1H), 6.98 (s, 1H), 5.68 (d, $J = 14.4$ Hz, 1H), 5.64 (d, $J = 14.4$ Hz, 1H), 4.54 (dd, $J = 7.8, 5.4$ Hz, 1H), 3.99 (s, 3H), 3.95 (s, 3H), 3.11 - 3.02 (m, 1H), 3.02 - 2.94 (m, 1H), 2.26 - 2.10 (m, 2H), 0.218 (s, 9H). ^{13}C NMR (151 MHz, CDCl_3) δ 170.2, 153.8, 148.6, 126.4, 119.4, 110.3, 108.4, 100.1, 66.0, 59.8, 56.6, 56.5, 36.2, 26.2, -0.3.

For compound **13**, compound **13a** (1.656 g, 3.864 mmol) was dissolved in a 1:1 mixture of 1,4-dioxane and concentrated HCl (8 mL). The reaction mixture was heated at 50 °C for 15 hours. The mixture was cooled to room temperature and diluted with distilled water. The reaction mixture was extracted 3X with CH_2CH_3 . The combined organic was dried over MgSO_4 then concentrated. The resulting yellow oil was purified by C18 column ($\text{H}_2\text{O}/\text{MeCN}$, 3:2) to yield compound **13** (0.8618 g, 59.4%).

HRMS analysis identified the composition: $\text{C}_{14}\text{H}_{17}\text{NO}_9\text{SNa}$. Measured mass was 398.0521; theoretical mass: 398.0516.

NMR analysis results for compound **13**: ^1H NMR (600 MHz, CDCl_3) δ 7.73 (s, 1H), 6.99 (s, 1H), 5.71- 5.64 (m, 2H), 4.41-4.30 (m, 1H), 4.00 (s, 3H), 3.96 (s, 3H), 3.15-3.03 (m, 2H), 2.29-2.19 (m, 1H), 2.13-2.03 (m, 1H).

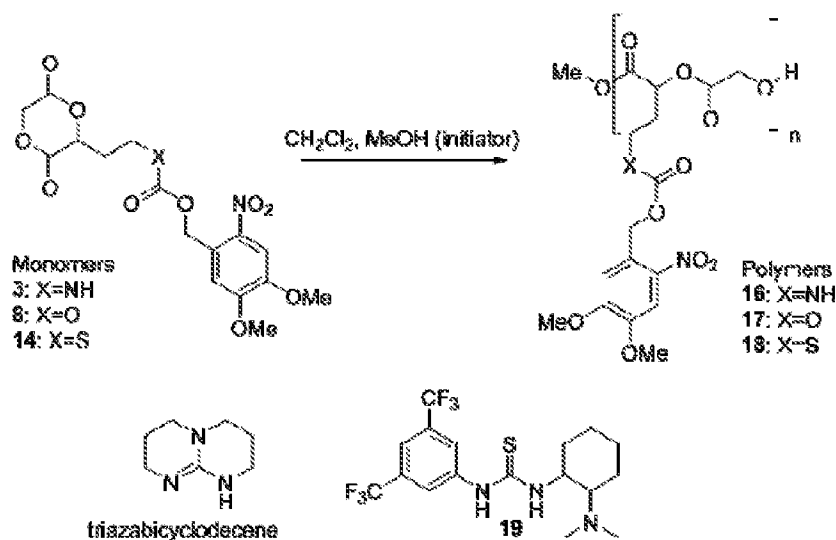
For compound **14**, compound **13** (0.823 g, 2.19 mmol) was dissolved in MeCN (14 mL) and chilled to 0 °C. Et_3N (0.398 mL, 2.85 mmol) after which a solution of bromoacetyl bromide (0.2489 mL, 2.85 mmol) in acetonitrile (7 mL) was dripped in over 10 min. The reaction mixture was diluted with EtOAc then washed 3X with IM HCl. The organic layer was dried over MgSO_4 and concentrated. The resulting orange oil was used without further purification. The oil was dissolved in DMF (33 mL) and dripped into a suspension of NaHCO_3 (276 mg, 3.29 mmol) in DMF (66 mL) over a period of 28 hours. The reaction was stirred for a further 8 hours. The reaction mixture was concentrated, then suspended in ethyl acetate. The suspension was filtered and the filtrate was collected and concentrated. The resulting oil was purified by silica column (1:1 Hex/ EtOAc) to yield compound **14** as a yellow solid (0.373 g, 41.0 %) over two steps.

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NMR analysis results for compound **14**: ^1H NMR (600 MHz, CDCl_3) δ 7.73 (s, 1H), 6.98 (s, 1H), 5.68 (d, $J = 14.4$ Hz, 1H), 5.64 (d, $J = 14.4$ Hz, 1H), 4.54 (dd, $J = 7.8, 5.4$ Hz, 1H), 3.99 (s, 3H), 3.95 (s, 3H), 3.11 - 3.02 (m, 1H), 3.02 - 2.94 (m, 1H), 2.26 - 2.10 (m, 2H), 0.218 (s, 9H). ^1H NMR (600 MHz, CDCl_3) δ 7.72 (s, 1H), 6.98 (s, 1H), 5.70 - 5.60 (m, 2H), 5.10 - 5.05 (m, 1H), 4.99 (d, $J = 16.8$ Hz, 1H), 4.94 - 4.88 (d, $J = 16.8$ Hz, 1H), 3.99 (s, 3H), 3.96 (s, 3H), 3.21 - 3.13 (m, 1H), 3.10 - 3.03 (m, 1H), 2.54 - 2.46 (m, 1H), 2.40 - 2.31 (m, 1H). ^{13}C NMR (151 MHz, CDCl_3) δ 170.1, 166.4, 164.1, 153.8, 148.7, 126.2, 110.7, 108.5, 100.1, 73.6, 66.21, 66.6, 56.7, 56.0, 31.1, 26.2.

Scheme 3 (also illustrated as FIG. 3) shows the preparation of the three light-degradable polymers **16**, **17**, and **18**: Triazabicyclodecene was used as catalyst and **19** was used as co-catalyst.

SCHEME 3



The three monomers were polymerized by organic-catalyzed ring-opening polymerization (ROP). We could not directly follow the method used by Dove et al. for this variety of monomer due to commercial unavailability of the catalyst (-)-sparteine. Instead, an alternative reported organocatalyst, triazabicyclodecene, was selected for ROP (Scheme 3). Commercially available compound **19**, 1-[3,5-bis(trifluoromethyl)phenyl]-3-[(1R,2R)-(-)-2-(dimethylamino) cyclohexyl]thiourea (R,R-TUC), was also included as a co-catalyst. The triazabicyclodecene alone was sufficient to catalyze ROP, but at a much slower rate. Attempts at polymerization with

metal catalysts were unsuccessful because of poor solubility in compatible solvents and the high melting points of the monomers in bulk. Weight-average molecular weights (MW) of the polymers were determined by gel permeation chromatography (GPC) relative to PMMA standards to be 3,800 Da (PDI = 1.2) for polymer **16**, 19,200 Da (PDI = 1.5) for polymer **17**, and 12,200 Da (PDI = 1.5) for polymer **18** using PMMA standards. Side reactions during ROP were a limiting factor in MW for the polymers. Polymer **16**'s length appeared to be limited by side reactions like transcarbamation.

The synthesis of polymers **16**, **17** and **18** was performed as follows:

10 Polymer **16**: Compound **3** (141 mg, 0.35 mmol) and compound **19** (26 mg, 0.063 mg) were suspended in CH₂Cl₂ (1.5 mL). 125 µL of 0.1 M methanol in CH₂Cl₂ were added dropwise to the reaction mixture. 125 µL of 0.1 M triazabicyclo[4.4.0]dec-5-ene in CH₂Cl₂ were added dropwise to the reaction mixture. The suspended compound **3** cleared as the reaction proceeded. The reaction
15 was allowed to proceed for 24 hours. The polymer was purified by repeated precipitation into cold hexanes from a CH₂Cl₂ solution to yield Polymer **16** (85 mg, 85%) as a yellow solid.

Polymer **17**: Compound **8** (141 mg, 0.35 mmol) and compound **19** (18 mg, 0.044 mmol) were dissolved in CH₂Cl₂ (1.5 mL). 89 µL of 0.1 M methanol in
20 CH₂Cl₂ were added dropwise to the reaction mixture. 89 µL of 0.1 M triazabicyclo[4.4.0]dec-5-ene in CH₂Cl₂ were added dropwise to the reaction mixture. The reaction was allowed to proceed for 15 hours. The polymer was purified by repeated precipitation into cold ether from a CH₂Cl₂ solution to yield Polymer **17** (141 mg, 100%) as a yellow solid.

25 Polymer **18**: Compound **14** (103 mg, 0.25 mmol) and compound **19** (13 mg, 0.031 mmol) were dissolved in CH₂Cl₂ (1.5 mL). 62 µL of 0.1 M methanol in CH₂Cl₂ were added dropwise to the reaction mixture. 62 µL of 0.1 M triazabicyclo[4.4.0]dec-5-ene in CH₂Cl₂ were added dropwise to the reaction mixture. The reaction was allowed to proceed for 20 hours. The polymer was
30 purified by repeated precipitation into cold ether from a CH₂Cl₂ solution to yield Polymer **18** (91 mg, 89%) as a yellow solid.

NMR analysis results were ^1H NMR (600 MHz, CDCl_3) δ 7.69 (s, 1H), 6.98 (s, 1H), 5.62 (s, 2H), 5.28 (s, 1H), 4.98 - 4.62 (m, 2H), 3.98 (s, 3H), 3.94 (s, 3H), 3.01 (s, 2H), 2.32 (s, 2H).

To characterize the polymer degradation, we first compared the sensitivities of the carbamate, carbonate, and thiocarbonate photocleavable protecting groups of the polymers **16**, **17**, and **18** respectively. Polymers were irradiated with UV light (1 mW/cm²) for the specified times up to 18 min and the change in absorbance at 346 nm was monitored. The peak at 346 nm, corresponding to the 4,5-dimethoxy-2-nitrobenzyl protecting group decreased, while a new peak at 400 nm, associated with the cleaved 4,5-dimethoxy-2-nitrosobenzaldehyde, formed (FIGs. 4A-4C). The percent absorbance was plotted over time (FIG. 4D). The three protecting groups are quite similar in sensitivity, though the carbonate protecting group of polymer **17** is slightly less sensitive than the other two.

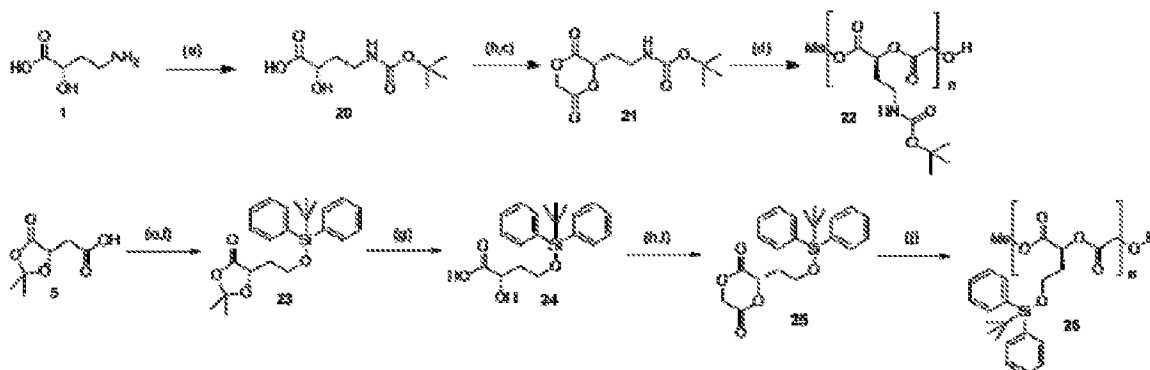
FIGs. 5A and 5B are GPC traces of a) polymer **17** and b) polymer **18** following 15 min irradiation (1 mW/cm²) and subsequent incubation for the specified times at 37 °C. The irradiation was relatively brief for this concentration so the cyclization chemistry could be readily observed.

Polymer degradation was monitored by gel permeation chromatography. Each polymer was dissolved in 9:1 acetonitrile/phosphate buffer pH 7.4 and irradiated for 15 min (1 mW/cm²), then incubated at 37 °C for specified times before concentrating the samples and analyzing by GPC. The 15 min of irradiation at this concentration is only enough to cleave a minor percentage of the protecting groups. This avoided substantial changes in the polarities of polymers **17** and **18**. Even with the low level of irradiation used polymer **16** underwent a substantial change in polarity, causing unfavorable interactions with the GPC columns that impeded interpretation (FIGs. 6A-C). Polymer **17** was amenable to GPC following irradiation; the initial irradiated trace (dashed line) shifted to longer elution times following incubation for 30 min (dotted line) and 1 hour (solid line) (FIG. 5A). This indicates a shift to lower molecular weight fragments, consistent with intramolecular cyclization. The minor difference between the 30 min. and 1 hr. traces suggests that the intramolecular cyclization reactions of polymer **17** are quite rapid, largely completing within 30 min of incubation. The molecular weight of polymer **18** did not change in 1 hr. (FIG. 5B). Changes did not

occur until after 4 hours of incubation, as can be seen in FIG. 7B. As this amount of time is compatible with hydrolysis of exposed thiocarbonate protecting groups (FIG. 5B), we infer that this structure did not cyclize appreciably at this temperature and the apparent degradation was unrelated to breaking of the polymer backbone. This result is not surprising, as the cyclization would necessitate an enthalpically unfavorable conversion of an ester to a thioester, though the reaction would be entropically favorable at higher (less biologically relevant) temperatures.

Scheme 4, which is also shown as FIG. 8, illustrates the synthesis of the model polymers using the following reagents and conditions: (a) Boc_2O , K_2CO_3 , H_2O , dioxane; (b) bromoacetyl bromide, Et_3N , MeCN, 0 °C; and (c) DMF, NaHCO_3 , 25% over two steps; (d) $\text{Sn}(\text{Oct})_2$, 125 °C, 29%. Synthesis of the silyl-protected model polymer; (e) borane, THF, 0 °C; and (f) TBDPSCl, pyridine, CH_2Cl_2 , 59% over two steps; (g) THF, H_2O , AcOH, 73%; (h) bromoacetyl bromide, Et_3N , MeCN, 0 °C; and (i) DMF, NaHCO_3 , 15% over two steps; (j) 19, MeOH, triazabicyclo[4.4.0]dec-5-ene, CH_2Cl_2 , 48%.

SCHEME 4



In Scheme 4, compounds **20-21** and Polymer **22** were synthesized as follows:

Compound 20: Compound **20** was synthesized from compound **1**, (*S*)-(-)-4-Amino-2-hydroxybutyric acid, via a previously described method.

Compound 21: Compound **20** (2.758 g, 12.59 mmol) was dissolved in acetonitrile (91 mL) and chilled to 0 °C. Et_3N (2.28 mL, 16.37 mmol) was dripped in followed by the dropwise addition of a solution of bromoacetyl bromide (1.43 mL, 16.37 mmol) in MeCN (30 mL). The reaction was stirred for 1.25 hour. The reaction was quenched with 1 M HCl, and extracted 3X with ethyl acetate. The

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combined organic was dried over MgSO_4 and concentrated. The resulting oil was used without further purification. The oil was dissolved in DMF (190 mL) and was dripped into a suspension of NaHCO_3 (1.586 g, 18.89 mmol) in DMF (381 mL) over a period of 28 hours. The reaction was stirred for a further 8 hours. The reaction mixture was concentrated then suspended in ethyl acetate. The suspension was filtered and the filtrate was collected and concentrated. The resulting oil was purified by silica column (1:1 Hex/EtOAc) then recrystallized (CH_2Cl_2 / Ether) to yield compound **20** as colorless crystals (0.808 g, 24.7%) over two steps.

HRMS analysis was used to determine the composition, which was: $\text{C}_{11}\text{H}_{17}\text{NO}_6\text{Na}$. Measured mass was 282.0946; theoretical mass: 282.0948.

NMR results for compound **21** were: ^1H NMR (600 MHz, CDCl_3) δ 5.03 (dd, $J = 7.8, 4.8$ Hz, 1H), 4.95 (d, $J = 16.2$ Hz, 1H), 4.91 (d, $J = 16.2$ Hz, 1H), 3.45-3.31 (m, 2H), 2.42-2.34 (m, 1H), 2.19-2.10 (m, 1H), 1.44 (s, 9H). ^{13}C NMR (151 MHz, CDCl_3) δ 165.9, 164.4, 156.2, 100.1, 79.9, 73.7, 36.3, 31.4, 28.5.

Polymer 22: Compound **21** (300 mg, 1.16 mmol) was placed in an oven dried flask. $\text{Sn}(\text{oct})_2$ solution 0.3 M in THF (8.56 μL) was dripped in and the mixture was heated to 125 $^\circ\text{C}$ where compound **22** melted and was allowed to react for 1.25 hour. The polymer was purified by repeated precipitation into cold ether from a DCM solution to yield Polymer **22** (85.9 mg, 28.6%) as a colorless solid.

The molecular weight of polymer **22** was determined to be 4,000 Da, PDI=1.2 by gel-permeation chromatography (GPC) relative to PMMA standards. NMR analysis results were ^1H NMR (600 MHz, CDCl_3) δ 5.25 (s, 1H), 5.05-4.66 (m, 3H), 3.43-3.18 (m, 2H), 2.23-2.01 (m, 2H), 1.43 (s, 9H).

Still referring to Scheme 4 (FIG. 8), compounds **23-25** and Polymer **26** were synthesized as follows:

Compound 23: Compound **5** (1.000 g, 5.7 mmol) was dissolved in THF (6 mL) in base washed glassware. The solution was chilled to 0 $^\circ\text{C}$ and a 1 M solution of borane in THF (8 mL) was added dropwise over 30 min. The reaction was allowed to warm to room temperature gradually and allowed to react for 3.5 hours. The reaction mixture was quenched with methanol at -78 $^\circ\text{C}$ and was concentrated. The reaction mixture was dissolved once more in methanol and concentrated then was dissolved in ethyl acetate and concentrated one more final time. The reaction

mixture was dissolved in CH_2Cl_2 (8 mL) without purification. The solution was chilled to 0 °C and pyridine (1.401 mL, 17.4 mmol) was dripped in. A solution of TBDPSCl (3.188 g, 11.6 mmol) in CH_2Cl_2 (12 mL) was slowly dripped into the reaction mixture. The mixture was allowed to warm to room temperature and react for 14 hours after which it was quenched with distilled water, organic collected and aqueous extracted 2X more with CH_2Cl_2 . Combined organic was washed with brine, dried over MgSO_4 , and concentrated. The concentrated mixture was purified by silica column (Hexane / 5% EtOAc) to yield compound **23** as colorless oil (1.364 g, 59.0%).

Using HRMS analysis, the composition of compound **23** was determined to be $\text{C}_{23}\text{H}_{30}\text{O}_4\text{SiNa}$. The measured mass was 421.1805; theoretical mass: 421.1806.

NMR analysis results were ^1H NMR (600 MHz, CDCl_3) δ 7.70-7.64 (m, 4H), 7.46-7.35 (m, 6H), 4.64 (dd, J = 8.4, 4.2 Hz, 1H), 3.92-3.87 (m, 1H), 3.82-3.75 (m, 1H), 2.20-2.12 (m, 1H), 1.96-1.88 (m, 1H), 1.59 (s, 3H), 1.55 (s, 3H), 1.05 (s, 9H). ^{13}C NMR (151 MHz, CDCl_3) δ 173.8, 135.7, 135.7, 133.7, 133.6, 129.8, 129.8, 127.8, 127.8, 110.7, 71.0, 59.4, 34.7, 27.4, 26.9, 25.9.

Compound **24**: Compound **23** (1.228 g, 3.1 mmol) was dissolved in 19 mL of THF, water, and AcOH 1:1:1. The reaction was stirred for 5 days. The reaction mixture was concentrated to yield compound **25** as a colorless oil (1.104 g, 73.2%) and was used directly to prepare compound **26**.

Using HRMS analysis, the composition of compound **24** was determined to be $\text{C}_{20}\text{H}_{25}\text{O}_4\text{Si}$. Measured mass was 357.1530; theoretical mass was 357.1528.

Compound **25**: Compound **24** (808 mg, 2.3 mmol) was dissolved in MeCN (16 mL) and chilled to 0 °C. Et_3N (0.41 mL, 2.9 mmol) was dripped in followed by the dropwise addition of a solution of bromoacetyl bromide (0.25 mL, 2.9 mmol) in acetonitrile (5 mL). The reaction was stirred for 1.25 hour. The reaction was quenched with 1 M HCl, and extracted 3X with ethyl acetate. The combined organic was dried over MgSO_4 and concentrated. The resulting oil was used without further purification. The oil was dissolved in DMF (34 mL) and was dripped into a suspension of NaHCO_3 (0.284 g, 3.4 mmol) in DMF (68 mL) over a period of 28 hours. The reaction was stirred for a further 8 hours. The reaction mixture was concentrated, then suspended in ethyl acetate. The suspension was filtered and the

filtrate was collected and concentrated. The resulting oil was purified by silica column (3:1 Hex/ EtOAc) to yield compound **25** as colorless oil (0.137 g, 15.3 %) over two steps.

NMR analysis results for compound **25** were ^1H NMR (600 MHz, CDCl_3) δ 7.71-7.61 (m, 4H), 7.49-7.36 (m, 6H), 5.13 (dd, $J = 8.4, 4.8$ Hz, 1H), 4.90 (d, $J = 16.8$ Hz, 1H), 4.85 (d, $J = 16.8$ Hz, 1H), 3.92-3.88 (m, 2H), 2.41-2.32 (m, 1H), 2.17-2.10 (m, 1H), 1.06 (s, 9H). ^{13}C NMR (151 MHz, CDCl_3) δ 166.1, 164.2, 135.6, 135.6, 133.2, 133.1, 130.1, 128.0, 72.4, 65.3, 58.4, 33.9, 27.0, 19.3.

Polymer **26**. Compound **25** (127 mg, 0.32 mmol) and compound **19** (17 mg, 0.04 mmol) were dissolved in CH_2Cl_2 (1 mL). 80 μL of 0.1 M methanol in CH_2Cl_2 was added dropwise to the reaction mixture. 80 μL of 0.1 M triazabicyclo[4.4.0]dec-5-ene in CH_2Cl_2 was added dropwise to the reaction mixture. The reaction was allowed to proceed for 15 hours. The polymer was purified by repeated precipitation into cold ether from a CH_2Cl_2 solution to yield Polymer **26** (62 mg, 48%) as a colorless solid.

The molecular weight of polymer **26** was determined to be 18,000 Da PDI= 1.7 by gel-permeation chromatography (GPC) relative PMMA. NMR analysis results for polymer **26** were ^1H NMR (600 MHz, CDCl_3) δ 7.66-7.58 (m, 4H), 7.42-7.30 (m, 6H), 5.49-5.41 (m, 1H), 4.84-4.70 (m, 1H), 4.64-4.50 (m, 1H), 3.81-3.65 (m, 2H), 2.34-2.23 (m, 1H), 2.11-1.95 (m, 1H), 1.05-0.97 (m, 9H).

We next sought to confirm that cyclizations driven by the pendant amine and alcohol nucleophiles contribute significantly to the degradation of polymers **16** and **17**. While identifying small cyclic degradation products would be the most direct means of confirming the mechanism, polymers **16** and **17** are not compatible with such an approach. As small products are only formed when two adjacent protecting groups are removed, generating sufficient quantities of such products for detection by NMR requires removal of most protecting groups. Such thorough deprotection would require intense and lengthy irradiation that causes numerous side reactions. To circumvent these problems, we synthesized two model polymers with conventional protecting groups that could be completely removed by chemical means (Scheme 4). Using model polymers **22** and **26** the degradation products of these backbones were studied and the presence of significant quantities of cyclic compounds was confirmed.

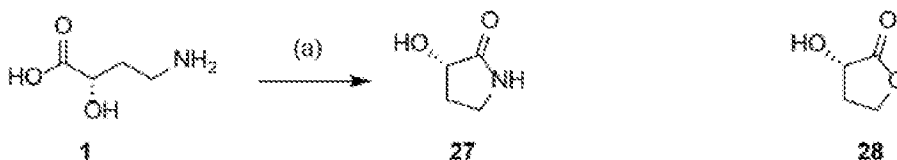
Model polymer **22**, analogous to polymer **16**, was prepared with a Boc in place of the light-sensitive protecting group, by a method similar to that used for polymer **16**. Boc protection of compound **1** yielded alcohol **20**, which was acylated with bromoacetyl bromide and subsequently cyclized with sodium bicarbonate to yield dilactone **21**. Monomer **21** proved difficult to polymerize, possibly due to an even stronger tendency towards transcarbamation than monomer **3**. Using Sn(Oct)₂ as a catalyst instead of the previously used organic catalyst provided low molecular weight polymer **22** (4,000 Da; PDI=1.2).

The model polymer analogous to polymer **17**, polymer **26**, was prepared incorporating a silyl protecting group for the alcohol. Using methods adapted from the synthesis of polymer **17**, compound **5** was reduced with borane, then protected with *tert*-butyl(chloro)diphenylsilane. The bulky protecting group was chosen to minimize deprotection of the silyl protecting group in the subsequent deprotection of ketal **23** with acetic acid to yield carboxylic acid **24**. Carboxylic acid **24** was readily converted into dilactone monomer **25**, using bromoacetyl bromide, which then was polymerized using the same technique as for the photosensitive polymers.

To confirm the presence of cyclic degradation products for polymer **22** and polymer **26**, they were deprotected with TFA or TBAF, respectively, and allowed to degrade. The deprotected materials were dissolved in deuterated buffer solutions prior to analysis by ¹H NMR spectroscopy. Minimal changes occurred in the spectra over time, indicating that degradation had reached near completion before samples could be analyzed by ¹H NMR spectroscopy, possibly during sample preparation or deprotection. Cyclic components were identified in the ¹H NMR spectra by comparison to predicted product spectra. Due to the nature of the ROP, the glycolic acid and the α-hydroxyl acid with a pendant nucleophile do not alternate perfectly, though certain patterns should be more likely due to steric interactions. This means that degradation could yield multiple cyclic products (FIGs. 9A, 9B). The two most readily synthetically accessible predicted products are compounds **27** and **28**. Compound **27** was prepared using a slightly modified method for a similar lactam, and compound **28** was readily obtained following a procedure developed by Denmark and Yang.

Scheme 5 shows the synthesis of model compound **27**.

SCHEME 5



Compound **27**: Compound **1** (1.0g, 8.40 mmol) and alumina (2.57g, 25.20 mmol) were suspended in toluene (43 mL) and heated to reflux with attached Dean-Stark apparatus. The reaction was allowed to proceed at reflux for 19 hours. The reaction was filtered and the solid washed with CH₂Cl₂. The filtrate was then collected and concentrated to yield compound **27** as colorless crystals (0.110 g, 12.9%).

Using HRMS analysis, the composition of compound **27** was determined to be $C_4H_7NO_2Na$. Measured mass was 124.0371; theoretical mass was 124.0369.

NMR analysis results for compound **27** were ¹H NMR (600 MHz, CDCl₃) δ 7.63 (bs, 1H), 5.35 (d, *J*= 5.4 Hz, 1H), 4.02-3.95 (m, 1H), 3.18-3.12 (m, 1H), 3.11-3.04 (m, 1H), 2.28-2.21 (m, 1H), 1.79-1.69 (m, 1H). ¹³C NMR (151 MHz, CDCl₃) δ 176.7, 68.4, 37.5, 30.5.

The peaks corresponding to these compounds were easily located in the degraded polymer spectra, confirming cyclization (FIGs. 9A, 9B). Peaks likely consistent with other cyclic products are also present in the degraded polymer spectra; the substantial difference between the methylene protons vicinal to the alcohol at 2-2.5 ppm is characteristic of those methylene protons when fixed in a ring. This evidence validates intramolecular cyclization as the major means of degradation for polymers of this backbone design upon deprotection.

FIG. 9A shows deprotection and subsequent degradation of model polymer **22** to likely cyclic degradation products and ^1H NMR spectra of degradation products of polymer **22** (upper trace) and compound **27** (lower trace). FIG. 9B shows deprotection and subsequent degradation of model polymer **26** to likely cyclic degradation products and ^1H NMR spectra of degradation products of polymer **26** (upper trace) and compound **28** (lower trace).

The three polymers were then formulated into nanoparticles to compare their degradation in a hydrophobic assembly and their potential for light-triggered release. Nanoparticles were formulated by single emulsion, both empty and encapsulating Nile red.

5 For formation of empty nanoparticles, 10 mg polymer was dissolved in 200 μ L chloroform, which was then added to 8 mL of 3% PVA in water. The mixture was sonicated using a 1/8 inch tip sonicator (Misonix S-4000) at about 9.5 W for 5 min. Particle solution was stirred at room temperature for 1 hour, followed by rotavap at 40°C to completely remove the organic solvent. The particle solution was then washed
10 with Hyclone water by tangential flow filtration through 500 kDa Pellicon XL cassettes (Millipore). 100 mg of Trehalose was added to the particle solution, which was then frozen by liquid nitrogen and lyophilized. Particles were formulated from polymers **16**, **17**, **18** and PLGA to yield nanoparticles NP **16**, NP **17**, NP **18** and PLGA particles respectively. FIGs. 10A-10C are TEM micrographs of nanoparticles NP **16**,
15 NP **17**, and NP **18**, respectively. Scale bar = 100 nm.

Nile red-loaded nanoparticles were formulated by dissolving 13 mg polymer in 200 μ L dichloromethane and combined with 100 μ L of Nile red solution in CH₂Cl₂ 0.13 mg/mL, which was then added to 6 mL of 1% PVA in water. The mixture was sonicated using a 1/8 inch tip sonicator (Misonix S-4000) at about 9.5 W for 5 min.
20 Particle solution was stirred at room temp for 2 hours under light vacuum to completely remove the organic solvent. The particle solution was then washed with Hyclone water by tangential flow filtration through 500 kDa Pellicon XL cassettes (Millipore). 120 mg of Trehalose was added to the particle solution, which was then frozen by liquid nitrogen and lyophilized. Particles were formulated from polymers
25 **16**, **17**, and **18** to yield nanoparticles NP-NR **16**, NP-NR **17**, and NP-NR **18** respectively.

To evaluate degradation, lyophilized particles were resuspended in IX PBS (pH 7.4) at 50 μ g/mL. 25 μ L of the particle solution was added to 1 mL of IX PBS (pH 7.4) in a
30 UV-transparent cuvette (dimension: 12.5 x 12.5 x 45 mm). The particle solution was irradiated with a band-pass filter (320 to 390 nm) at 0.181 W/cm² for 15 sec. DLS

was used to monitor the size, PDI and count rate continuously immediately after irradiation.

Degradation of polymers **16**, **17**, and **18** measured was by GPC. The GPC traces for each polymer are shown in FIGs. 6A-C, 7A-C. The polymer was dissolved in a mixture of acetonitrile and PBS (1X) 90:10 at a 0.2 mg/mL concentration. Irradiated samples were irradiated in a Luzchem photoreactor for 15 min (1 mW/cm²). The samples were incubated at 37 °C for the specified times, i.e., 30 min., 1 hour, 4 hours, 23 hours, and 48 hours. At the given times, the samples were concentrated at 30 °C, dissolved in DMF with 0.01 % LiBr, and analyzed by gel permeation chromatography monitoring at 320 nm.

The degradation of empty particles was continuously monitored by DLS following irradiation for 15 sec (0.181 W/cm²) (FIG. 11B). The count rate for irradiated particles of polymer **16** decreased more rapidly than for other particles. This rapid degradation likely results from the major increase in hydrophilicity upon release of the amine, as well as the amine's high nucleophilicity allowing rapid intramolecular cyclization. Particles of polymer **17**, NP **17**, also rapidly degrade. The count rate of irradiated particles of polymer **18** decreased at a rate similar to non-irradiated NP **18** (FIG. 11C). Particles formulated with PLGA behaved identically when irradiated and not irradiated under these conditions (FIG. 11D). To confirm NP degradation, transmission electron microscopy (TEM) micrographs of irradiated and non-irradiated particles were also obtained. In agreement with the DLS data, particle densities of NP **16** and NP **17** were markedly lower after irradiation 15 sec (0.181 W/cm²) and overnight incubation (e.g., 24 and 48 hours) at 37 °C. After irradiation, NP **16** (FIG. 11A) and NP **17** (FIGs. 12A, 12B) also contained substantial quantities of aggregates with no clear spherical structure, likely material from degraded particles. Irradiated NP **18** (FIGs. 12C, 12D) still had substantial numbers of intact particles, though the particles did appear to exhibit increased aggregation following irradiation.

To assess utility for triggered release, particles were also formulated encapsulating the fluorescent dye Nile red (NP-NR). Nile red is fluorescent in hydrophobic environments, like that of a hydrophobic nanoparticle, and is quenched by water. FIGs. 13A-13C are SEM images of particles encapsulating Nile red, where FIG. 13A shows NP **16**, FIG. 13B shows NP **17**, and FIG. 13C shows NP **18**

formulated by single emulsion. Upon irradiation with UV light (1 mW/cm^2), Nile red fluorescence was rapidly quenched in all three particles, but most quickly in NP-NR **16** and NP-NR **17**, which are composed of the polymers that degrade rapidly through cyclization (FIG. 14A). Rapid quenching is indicative of substantial structural changes to the particles, allowing both release of Nile red and water influx. Particles are stable over at least four hours when not irradiated (FIG. 14B).

Disclosed herein are novel polymers with a poly(lactide-*co*-glycolide)-type backbone that contain pendant protected nucleophiles. In polymers containing pendant alcohols or amines, polymer degradation following deprotection is accelerated by intramolecular cyclization events that cause breaks in the polymer backbone. This backbone has the potential to support a variety of protecting groups sensitive to different triggers.

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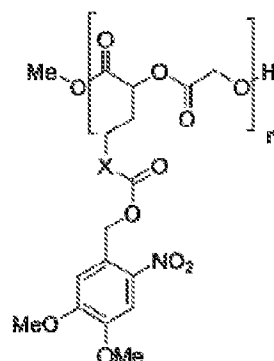
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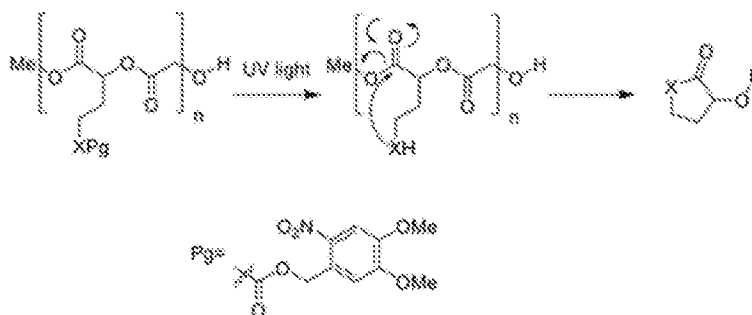
CLAIMS:

1. A composition comprising a polymer having a poly(lactide-*co*-glycolide)-type backbone and pendant nucleophiles protected by a stimulus-responsive protecting group, wherein the protecting group is configured to deprotect upon exposure to a stimulus to facilitate degradation by intramolecular cyclization.
2. The composition of claim 1, wherein the protecting group is photocleavable and the stimulus is light.
3. The composition of claim 2, wherein the light is UV light.
4. The composition of claim 1, wherein the protecting group is an ortho-nitrobenzyl (ONB) protecting group.
5. The composition of claim 1, wherein the polymer comprises:



or a derivative thereof, where X is NH, O or S.

6. The composition of claim 1, wherein the polymer degrades according to:



where X is NH, O or S.

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7. The composition of claim 1, wherein the polymer comprises a nanoparticle, and further comprising a payload encapsulated within the nanoparticle.

8. The composition of claim 7, wherein the payload comprises an imaging agent.

5 9. The composition of claim 7, wherein the payload comprises a bioactive agent.

10. The composition of claim 7, wherein the payload comprises a pharmaceutical agent.

11. The composition of claim 1, wherein the pendant nucleophile is an amine.

10 12. The composition of claim 1, wherein the pendant nucleophile is an alcohol.

13. The composition of claim 1, wherein the pendant nucleophile is a thiol.

14. A composition comprising a polymer comprising:

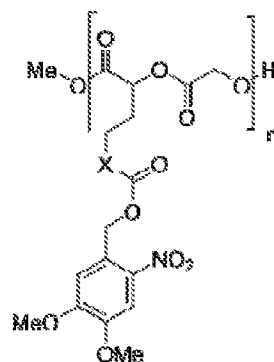
a polymer backbone comprising a poly(lactide-*co*-glycolide);

15 pendant nucleophiles linked to the backbone by photolabile protecting group, wherein the protecting group is configured to deprotect upon exposure to irradiation to facilitate degradation of the nucleophiles to five membered rings by intramolecular cyclization.

15. The composition of claim 14, wherein the irradiation is UV light.

20 16. The composition of claim 14, wherein the protecting group is an ortho-nitrobenzyl (ONB) protecting group.

17. The composition of claim 14, wherein the polymer comprises:

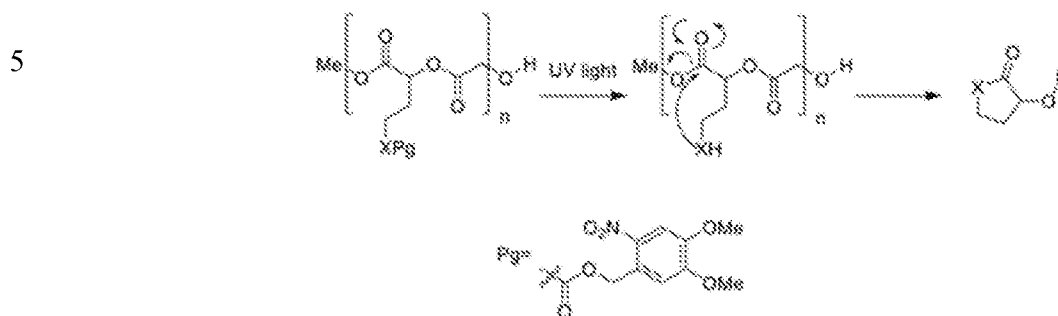


25

-34-

or a derivative thereof, where X is NH, O or S.

18. The composition of claim 14, wherein the polymer degrades according to:



where X is NH, O or S.

19. The composition of claim 14, wherein the polymer comprises a nanoparticle, and further comprising a payload encapsulated within the nanoparticle.

20. The composition of claim 19, wherein the payload comprises an imaging agent.

21. The composition of claim 19, wherein the payload comprises a bioactive agent.

22. The composition of claim 19, wherein the payload comprises a pharmaceutical agent.

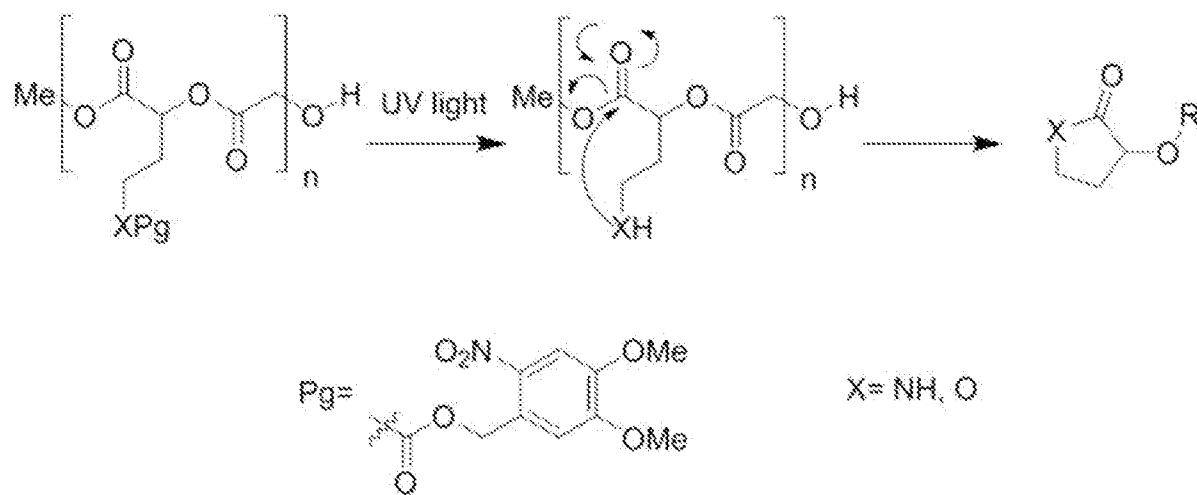
23. The composition of claim 14, wherein the pendant nucleophile is an amine.

24. The composition of claim 14, wherein the pendant nucleophile is an alcohol.

25. The composition of claim 14, wherein the pendant nucleophile is a thiol.

25

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SCHEME 1

FIG. 1A

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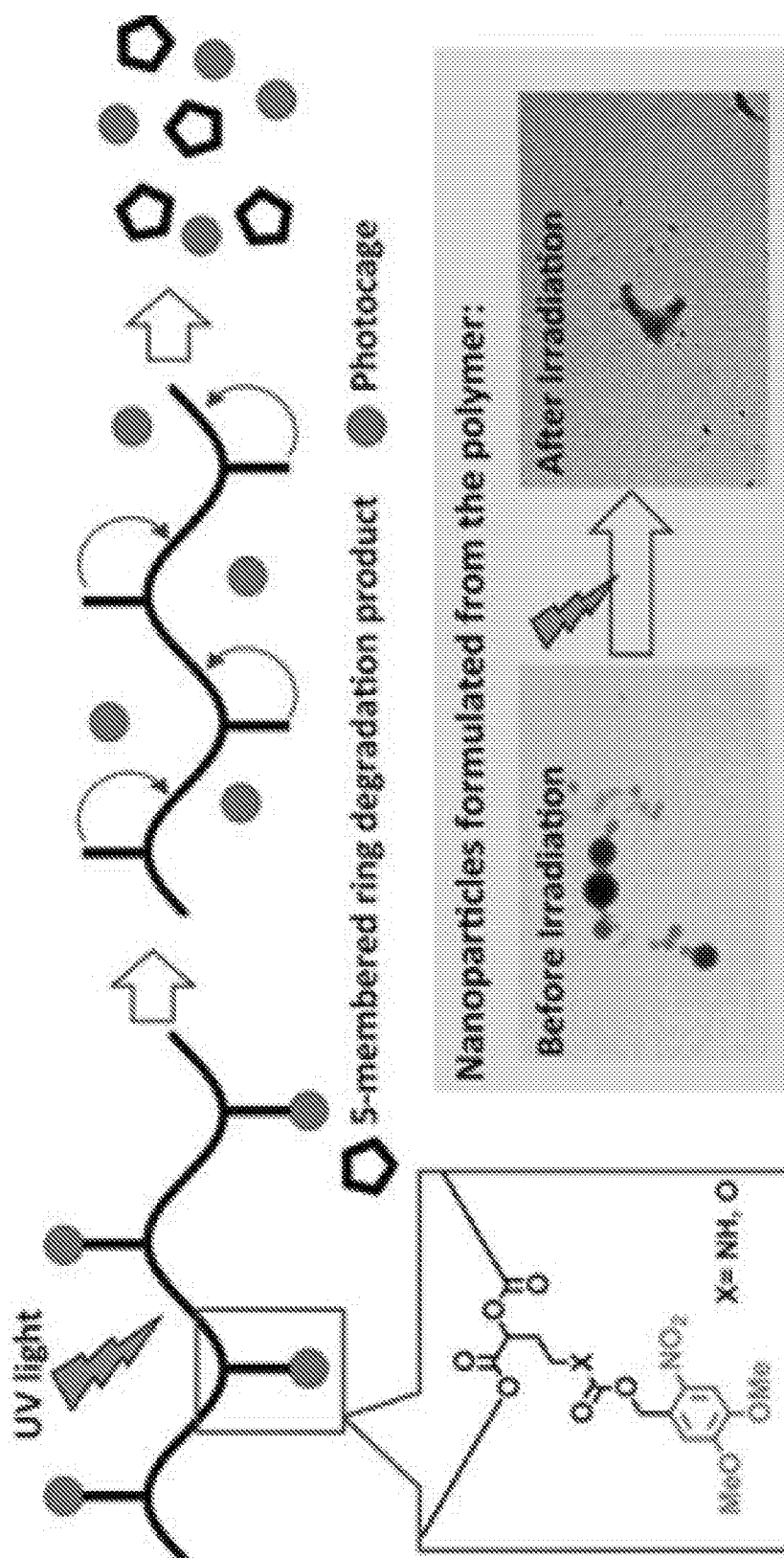


FIG. 1B

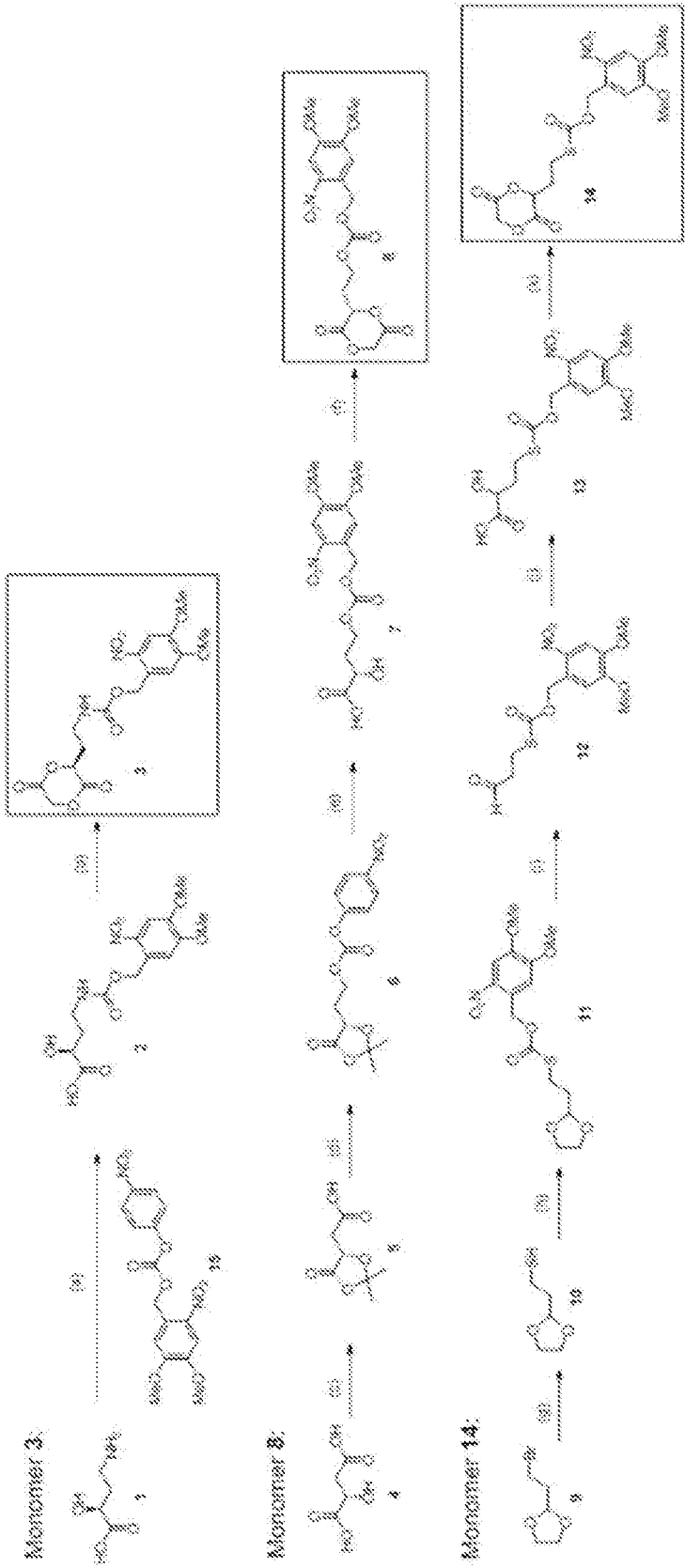


FIG. 2A

SCHEME 2

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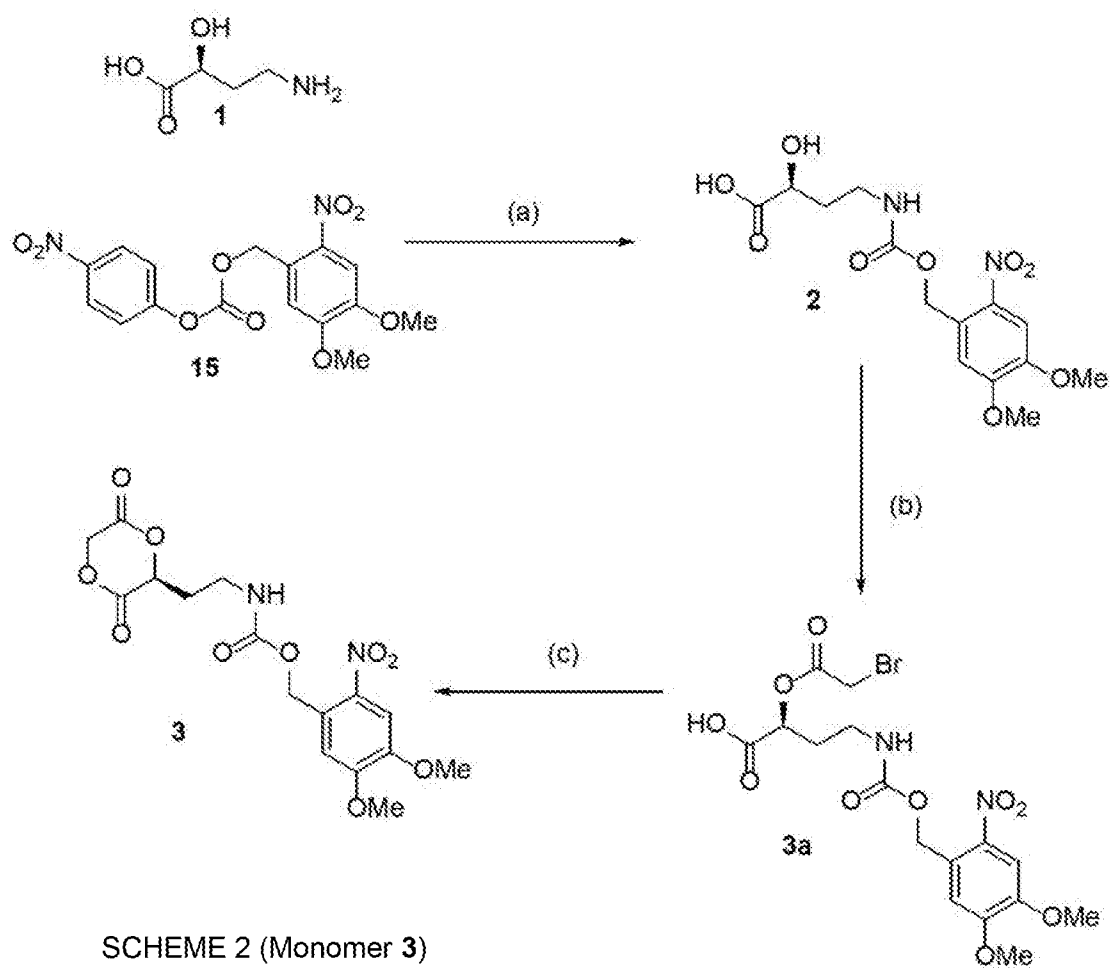


FIG. 2B

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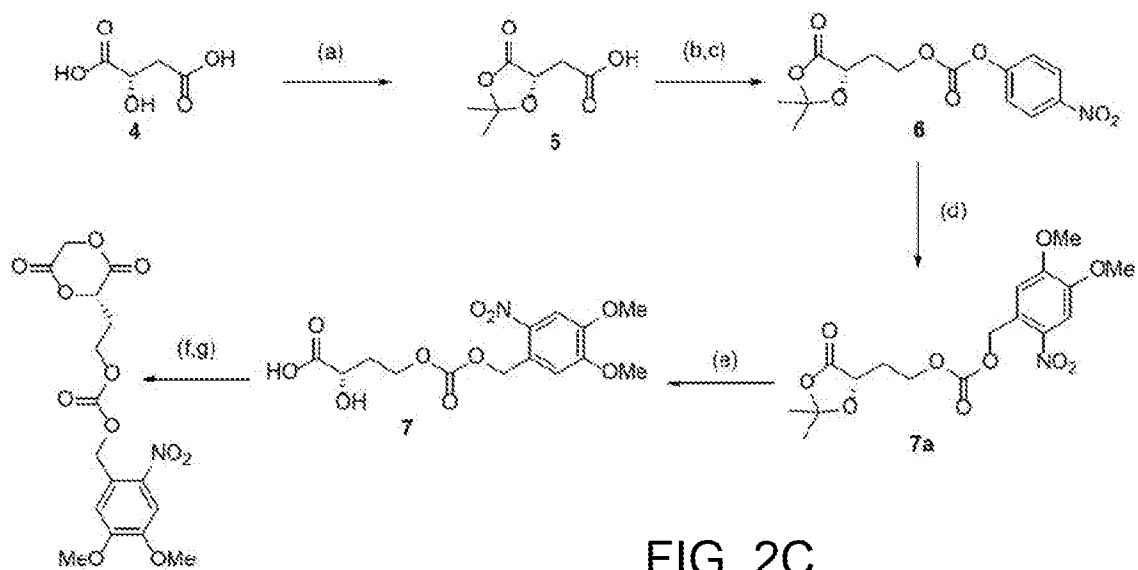
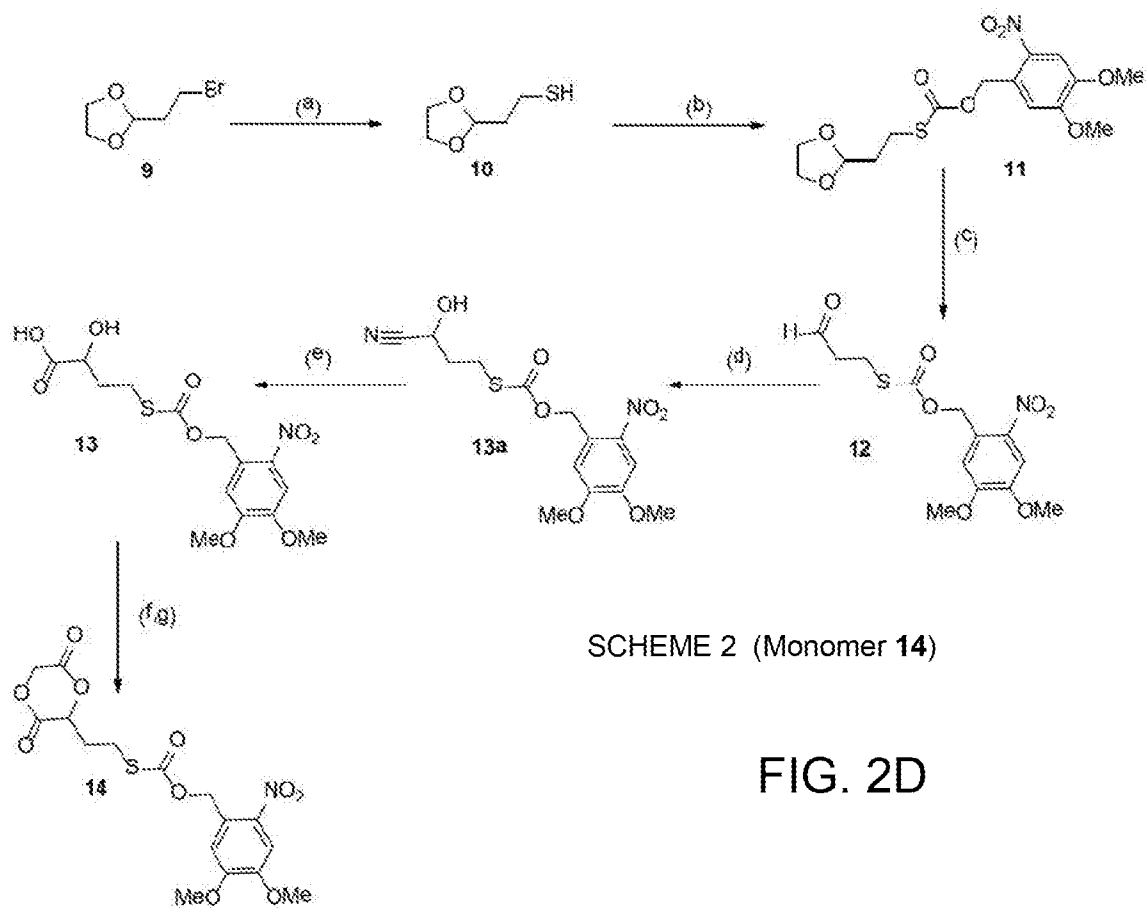


FIG. 2C

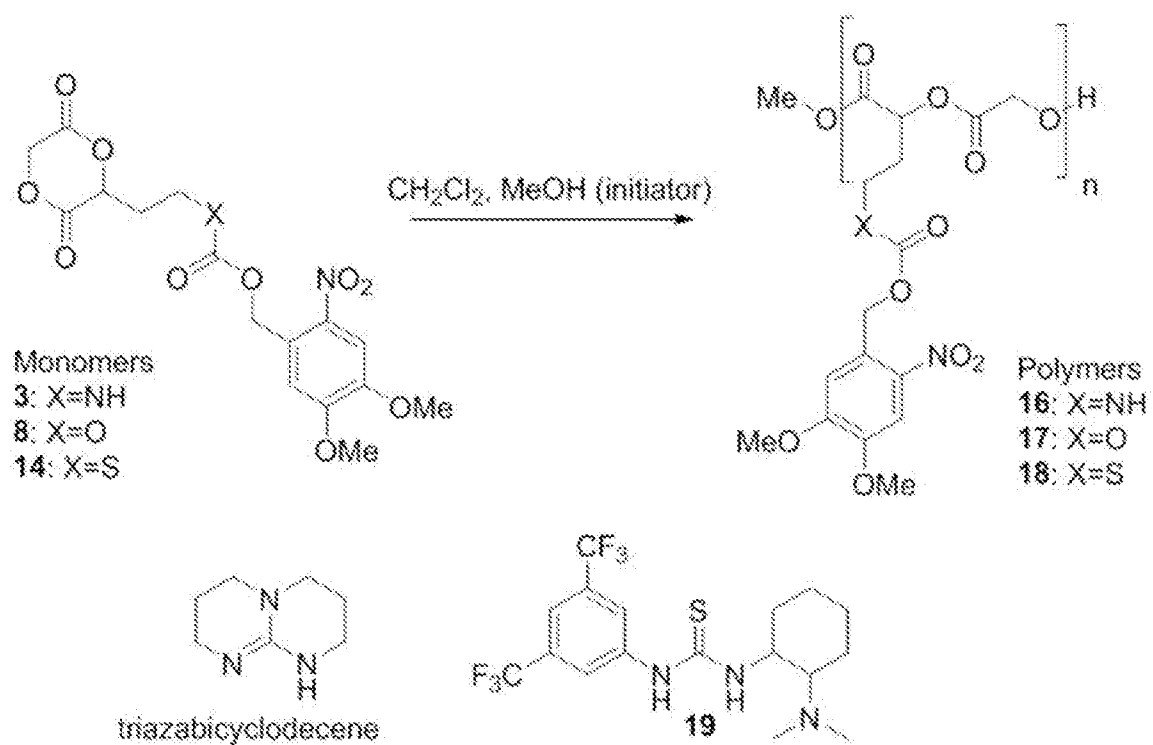
SCHEME 2 (Monomer 8)



SCHEME 2 (Monomer 14)

FIG. 2D

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SCHEME 3

FIG. 3

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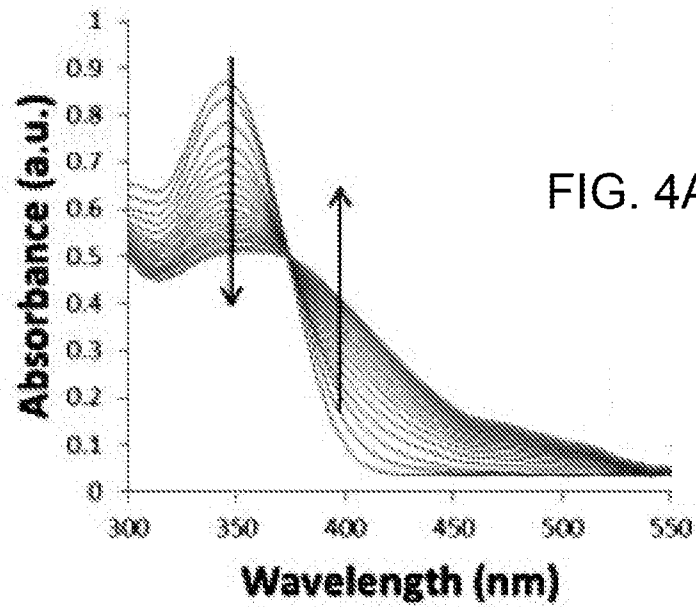


FIG. 4A

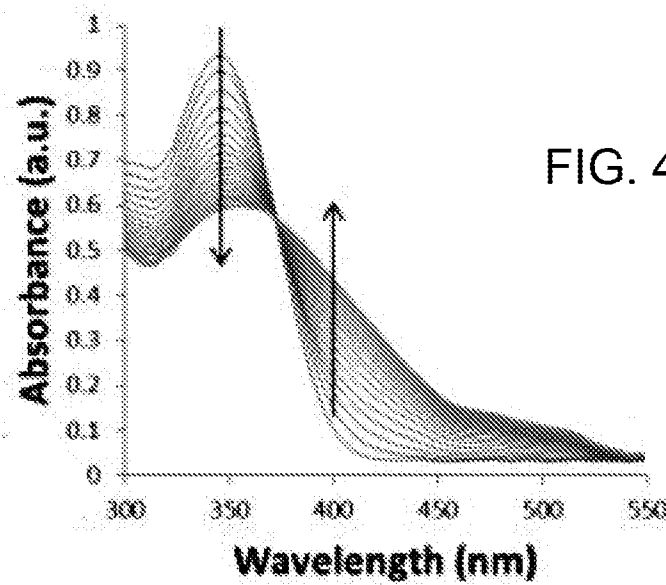


FIG. 4B

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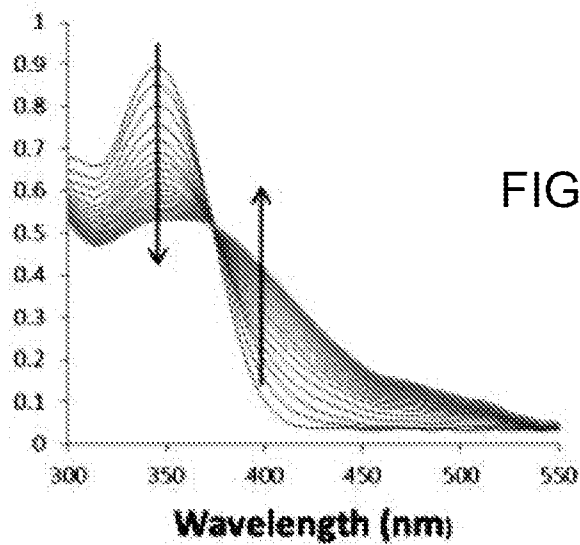


FIG. 4C

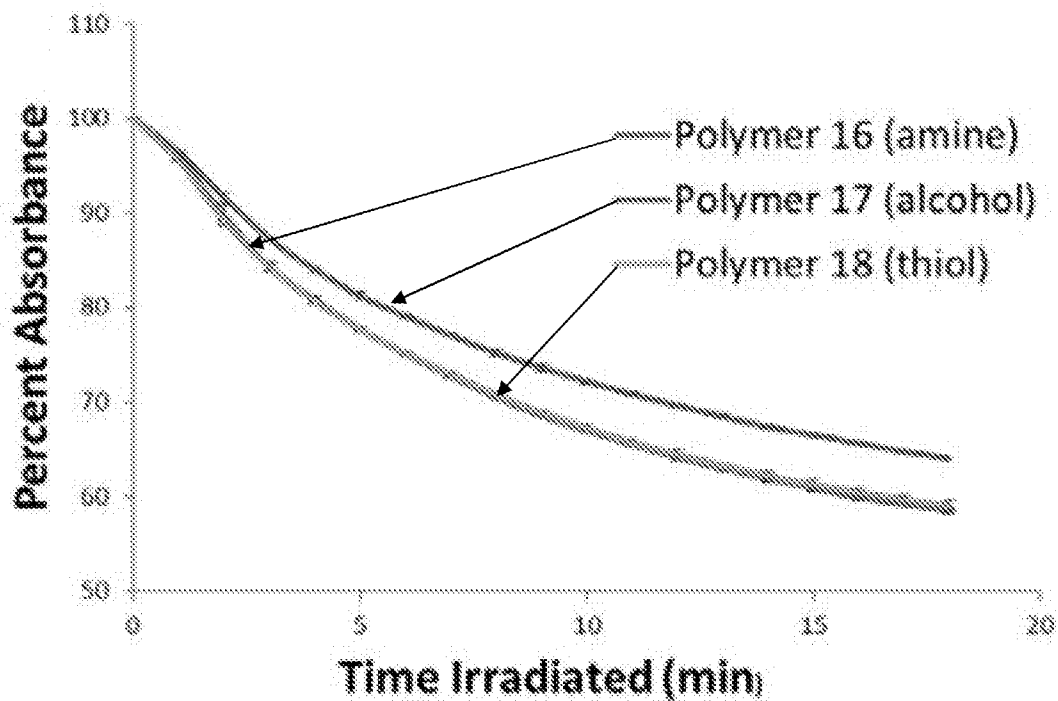


FIG. 4D

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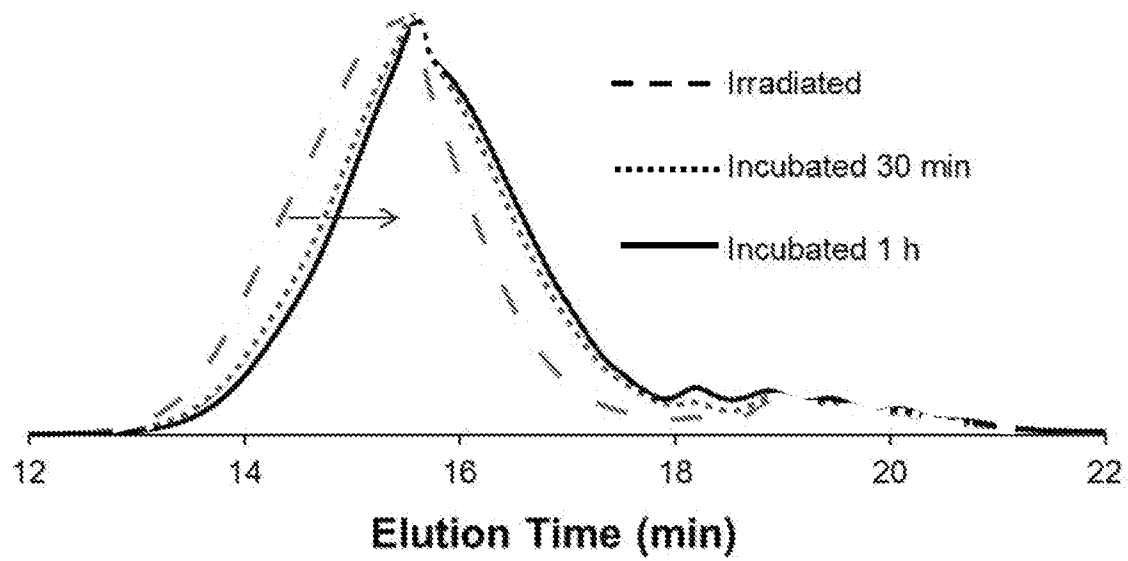


FIG. 5A

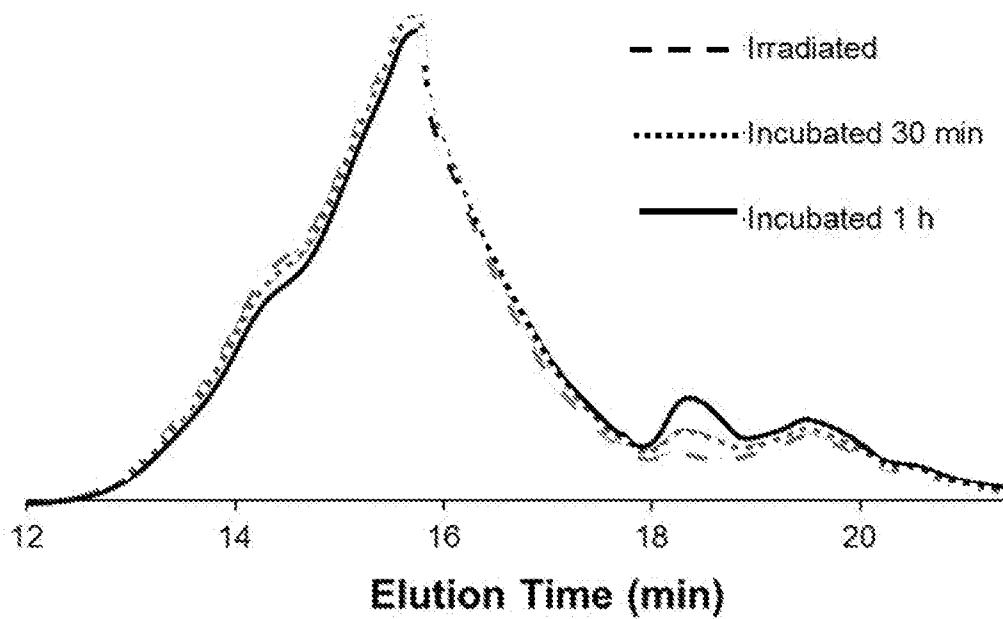


FIG. 5B

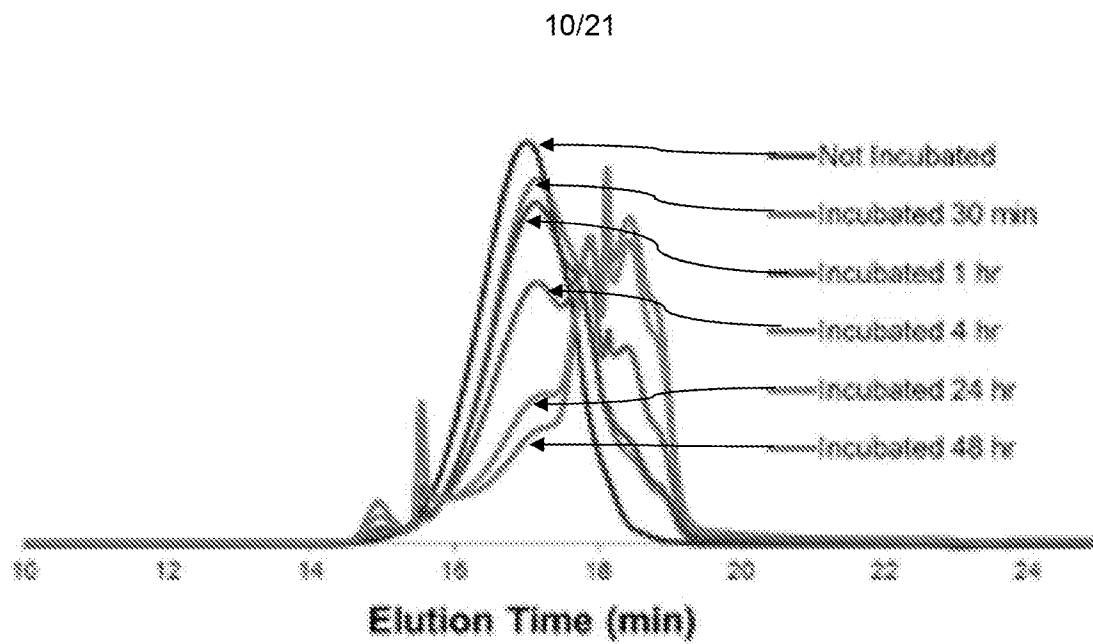


FIG. 6A

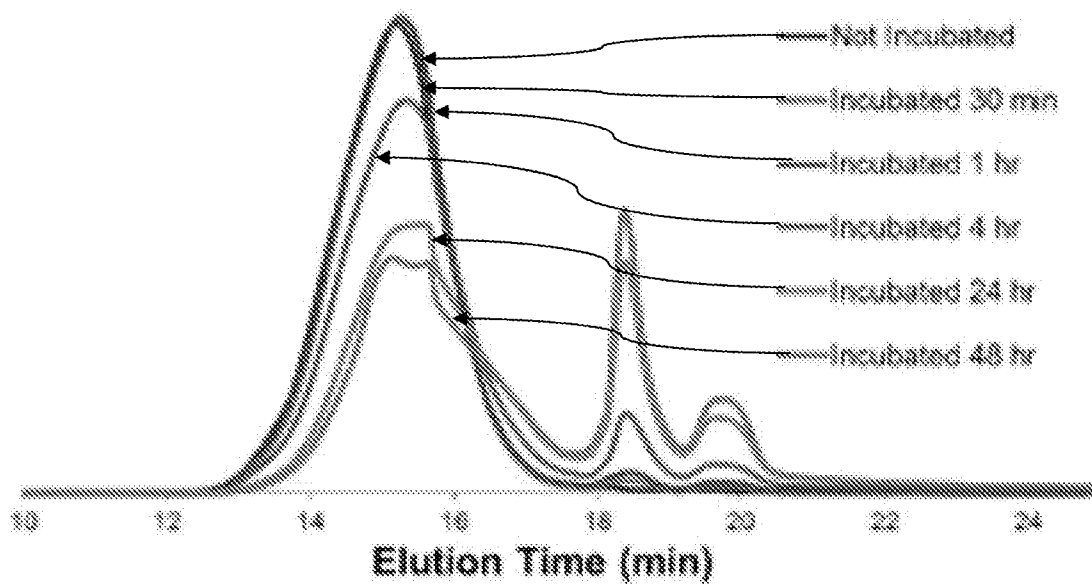


FIG. 6B

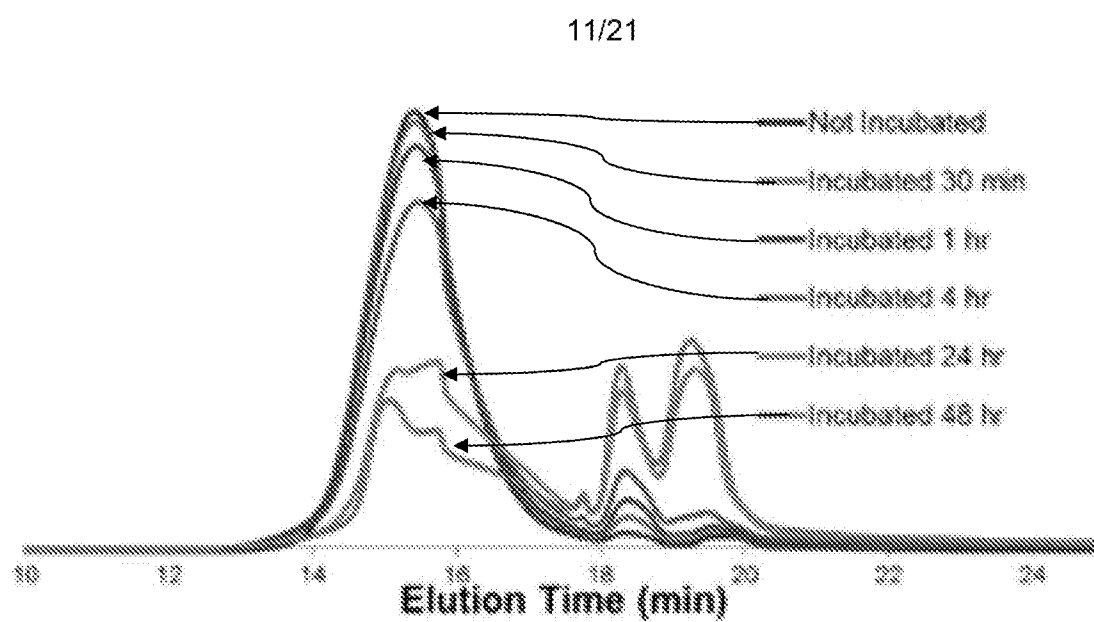


FIG. 6C

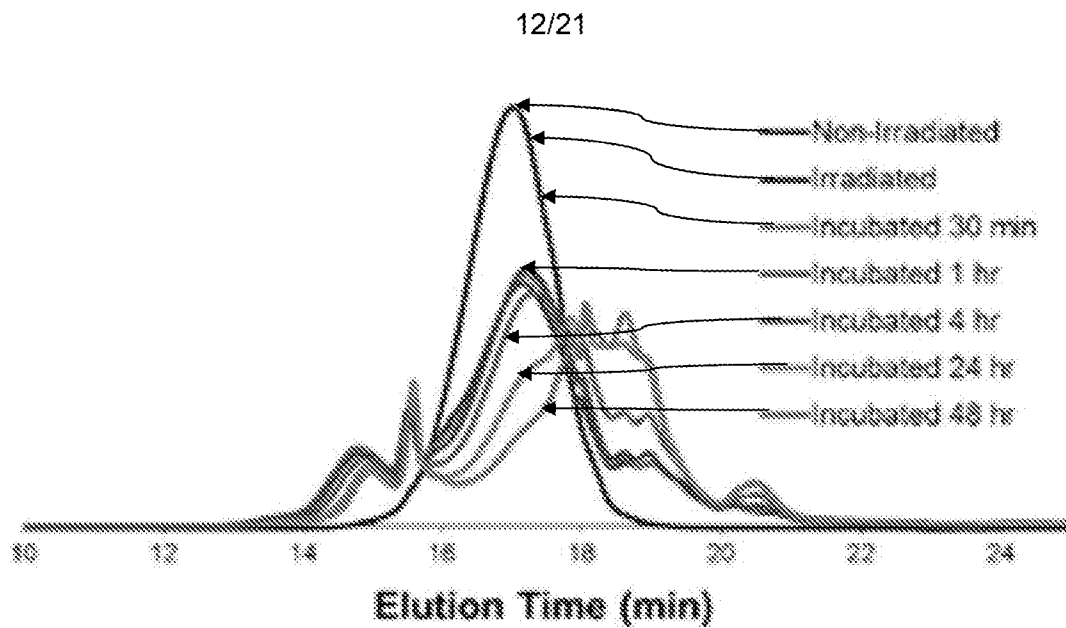


FIG. 7A

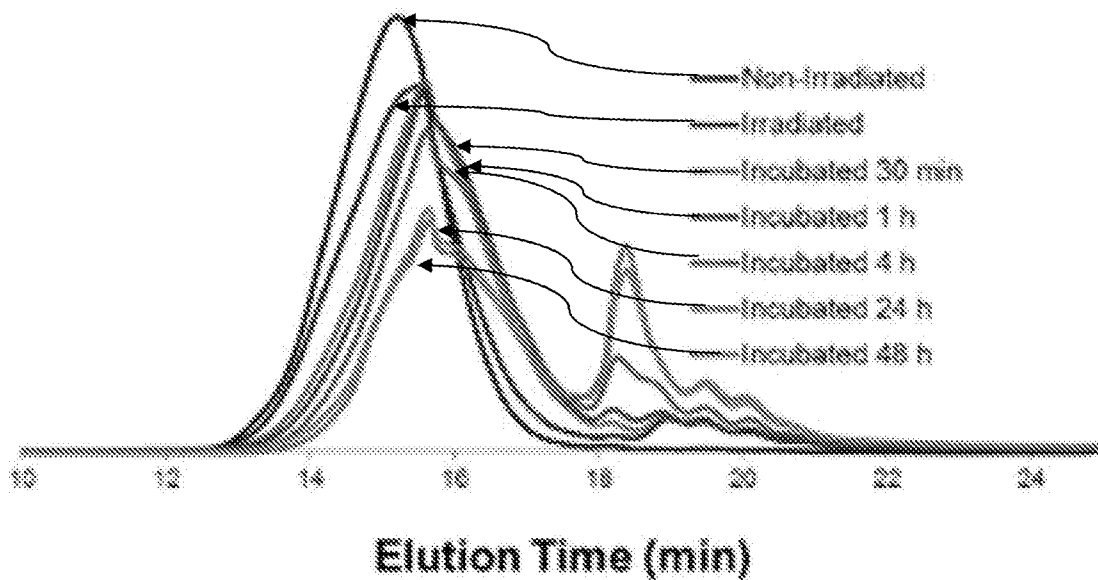


FIG. 7B

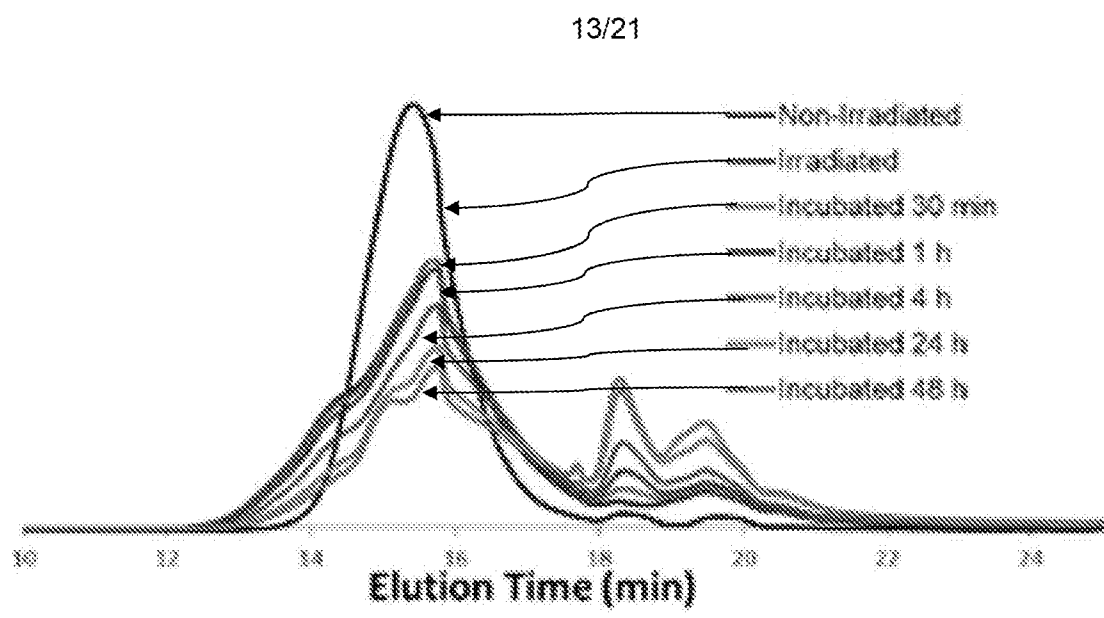
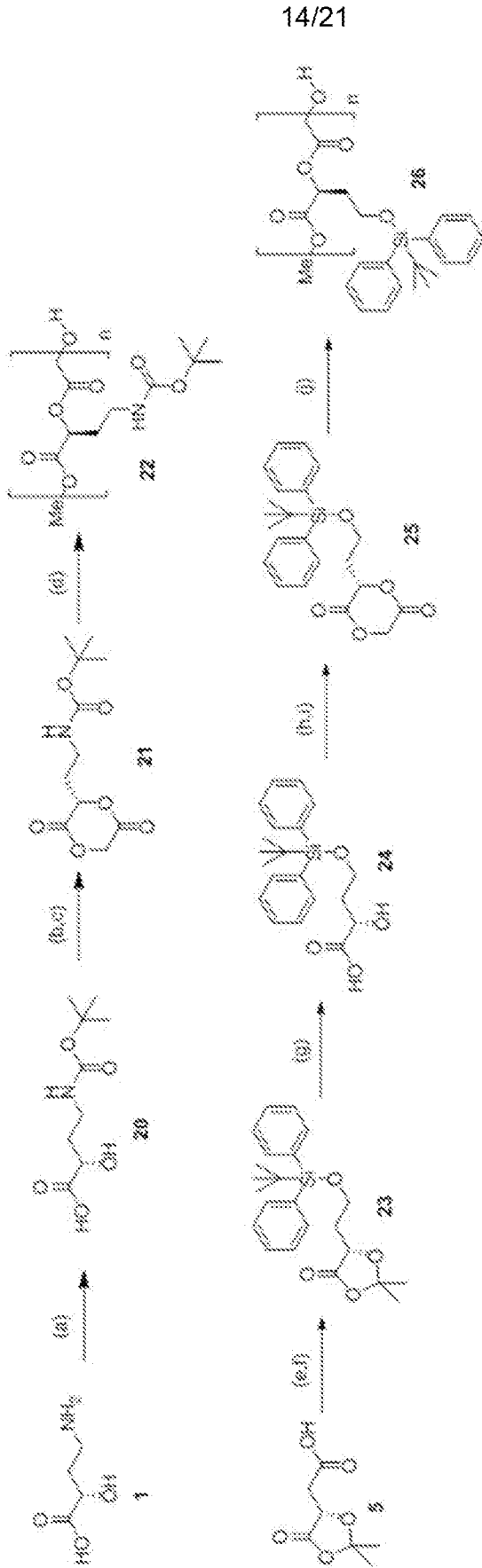


FIG. 7C



SCHEME 4

FIG. 8

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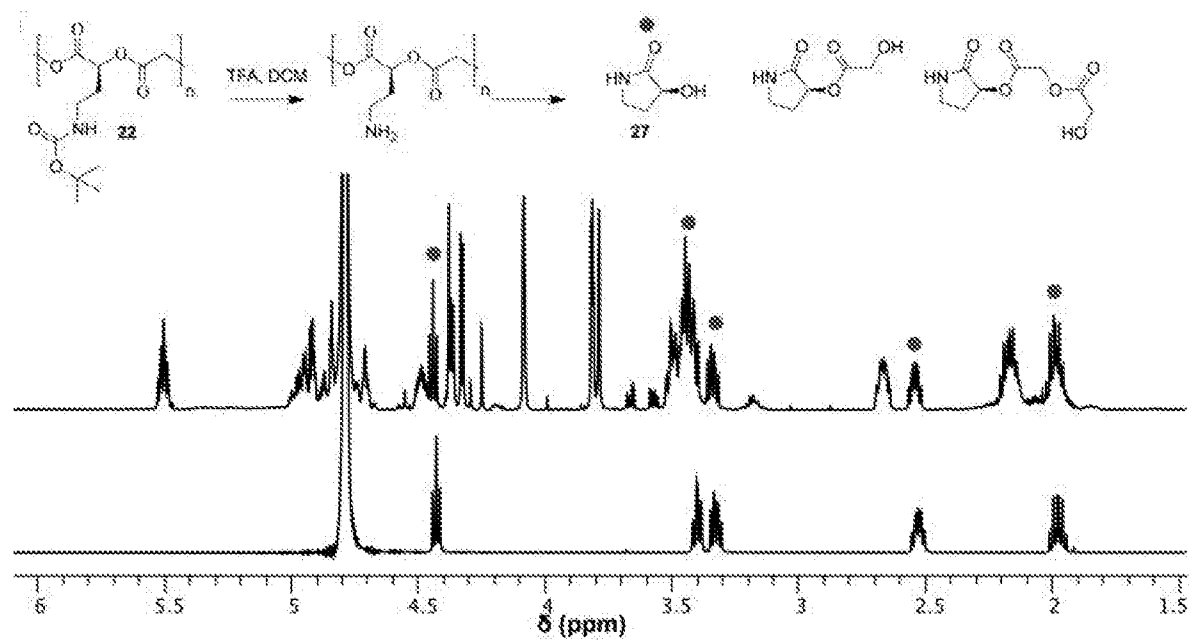


FIG. 9A

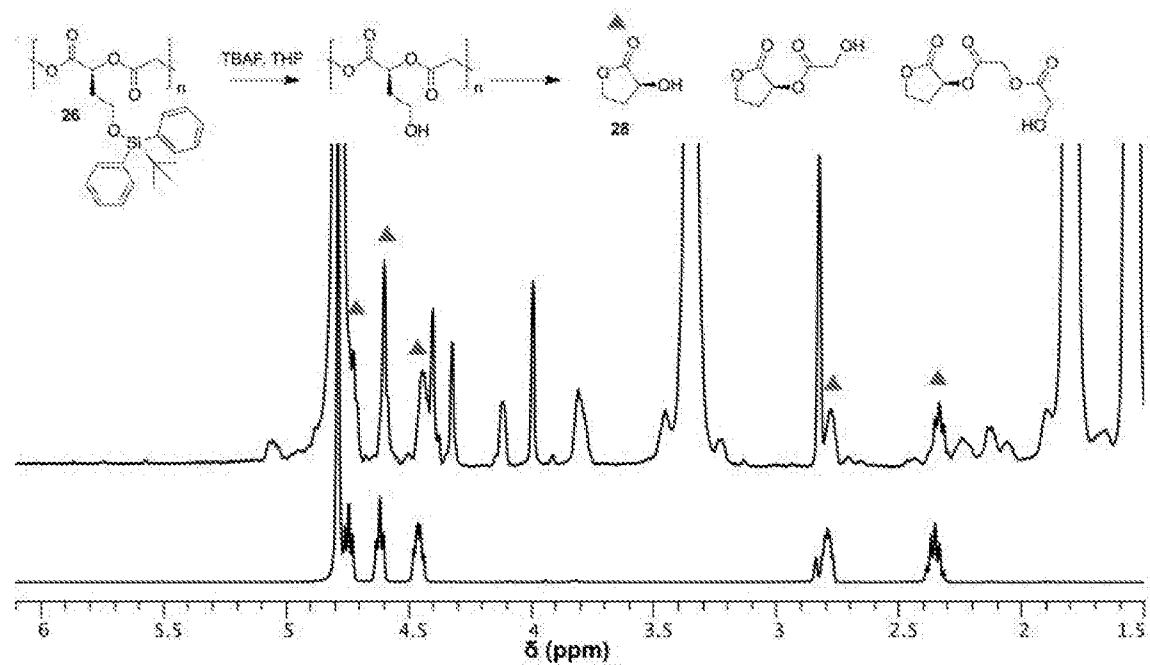


FIG. 9B

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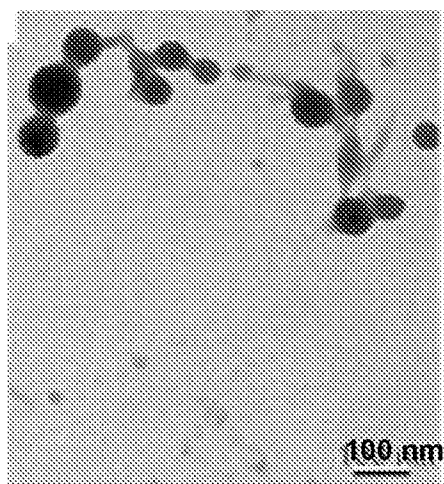


FIG. 10A

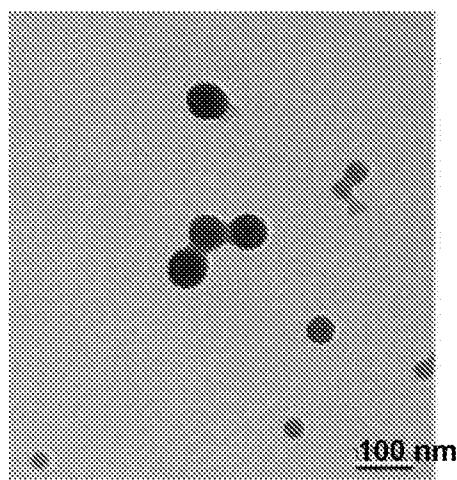


FIG. 10B

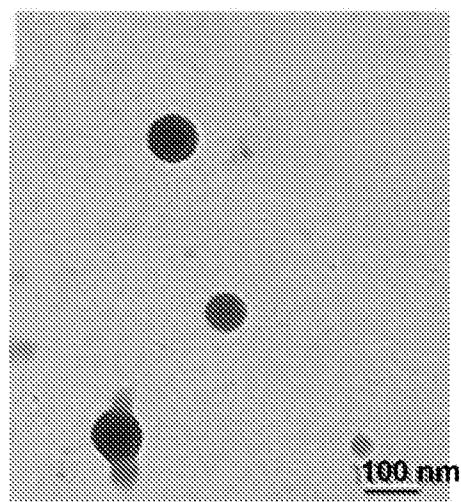


FIG. 10C

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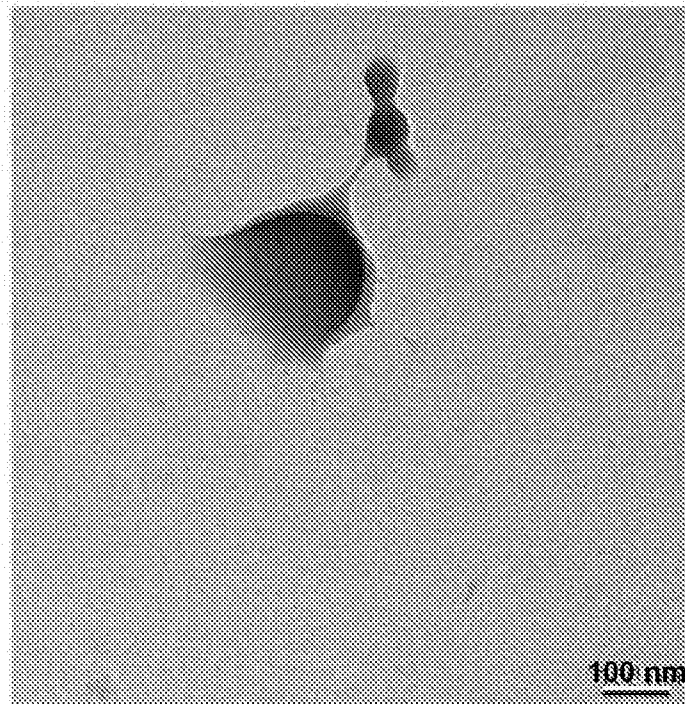


FIG. 11A

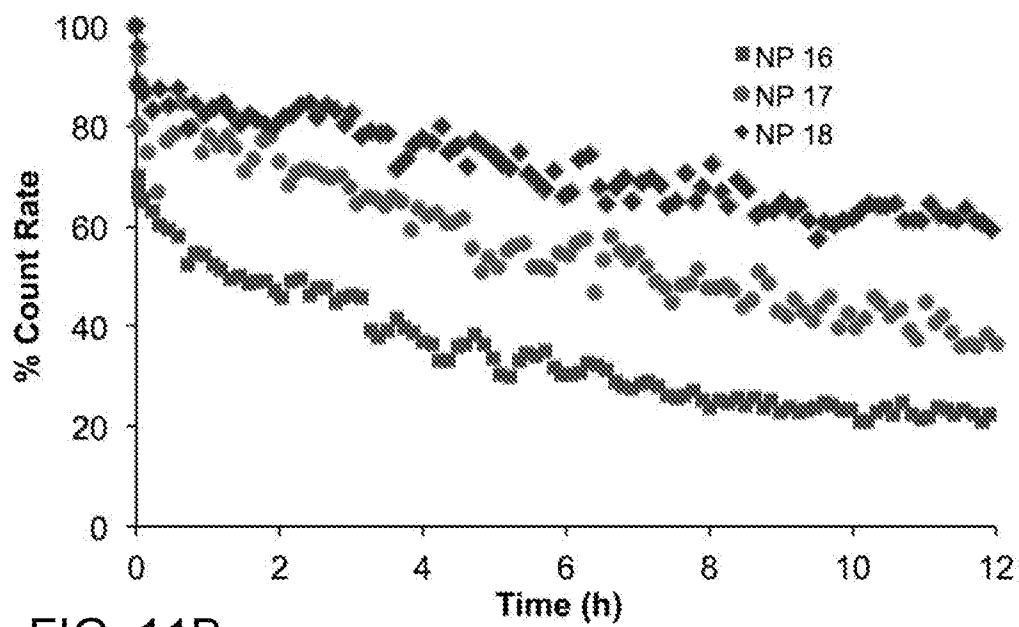


FIG. 11B

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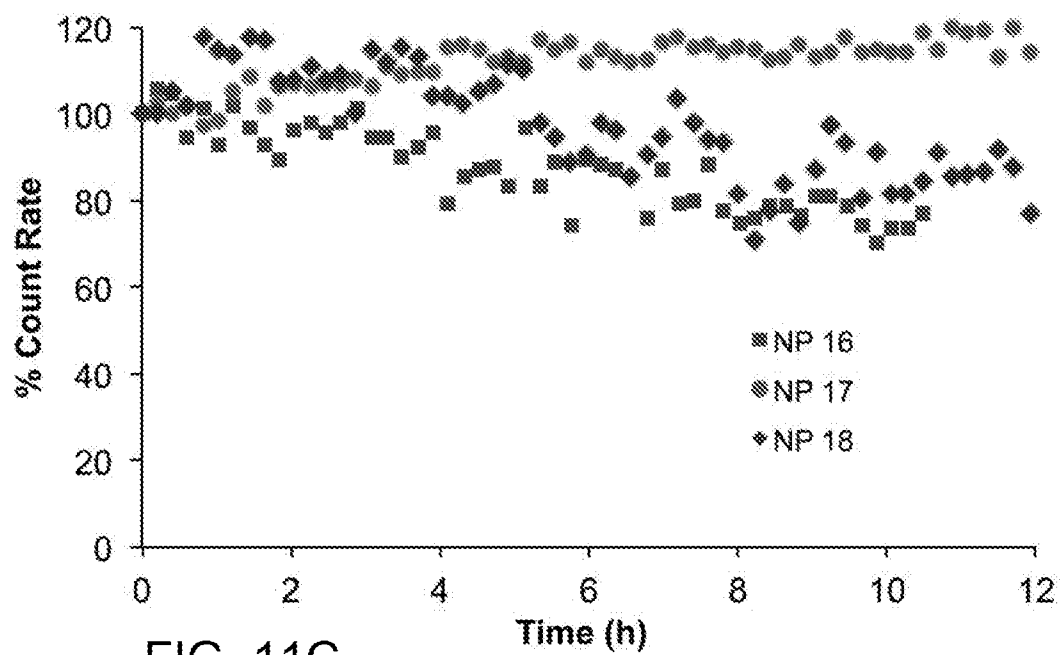


FIG. 11C

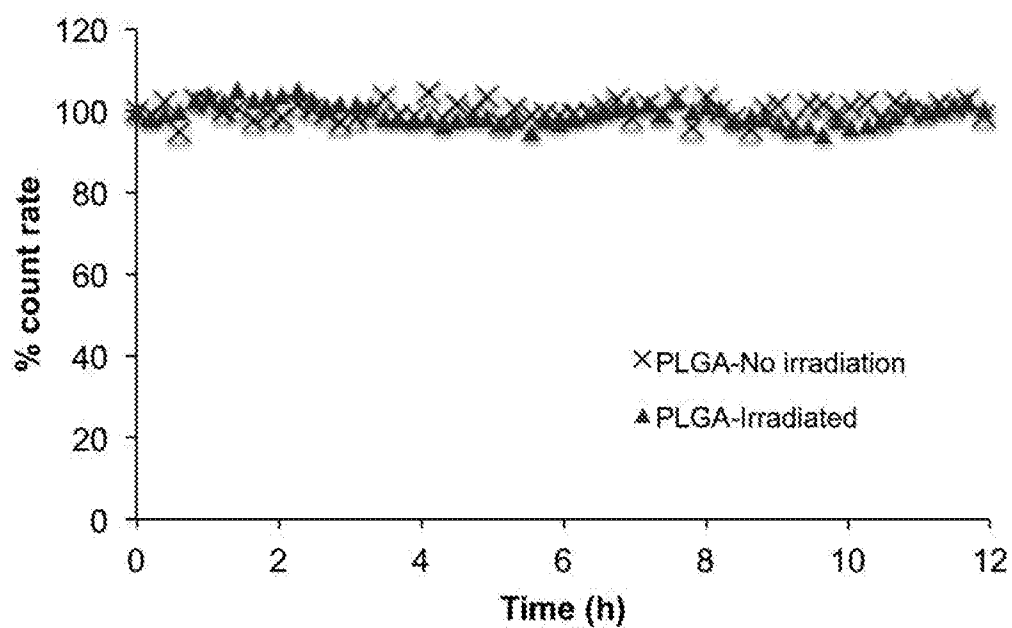


FIG. 11D

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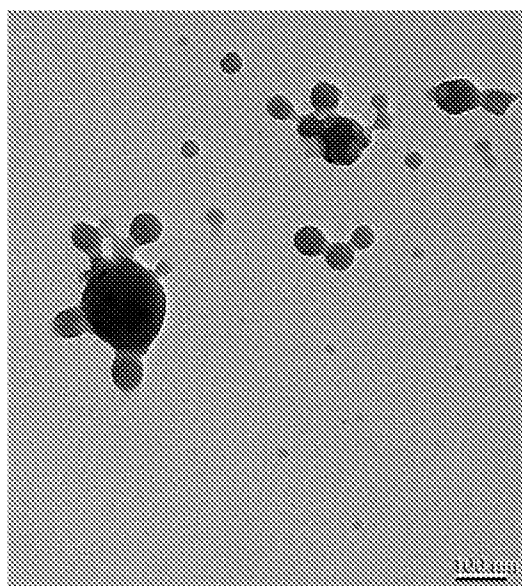


FIG. 12A

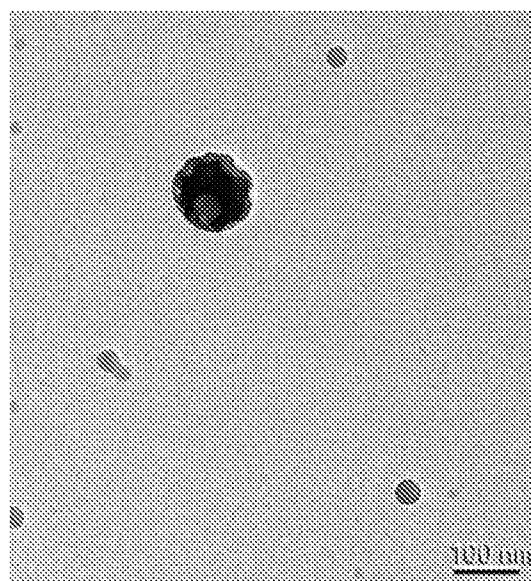


FIG. 12B

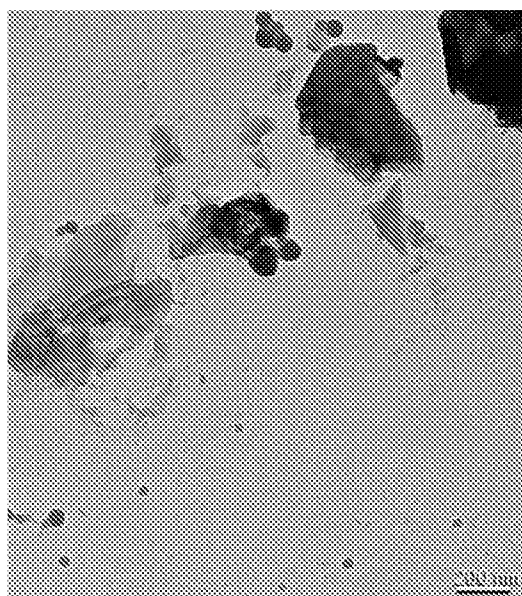


FIG. 12C

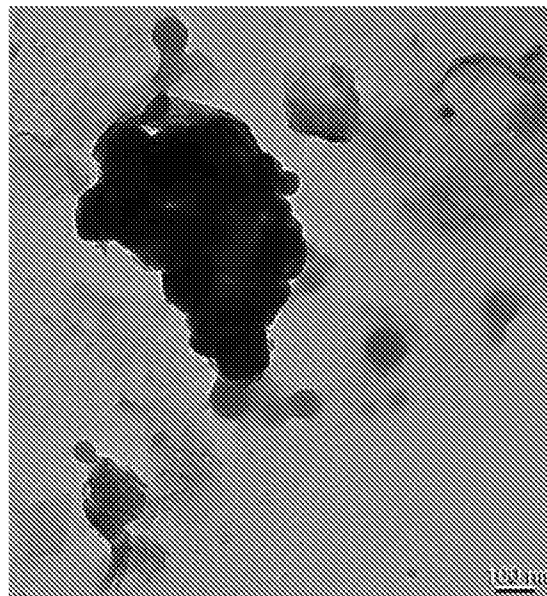


FIG. 12D

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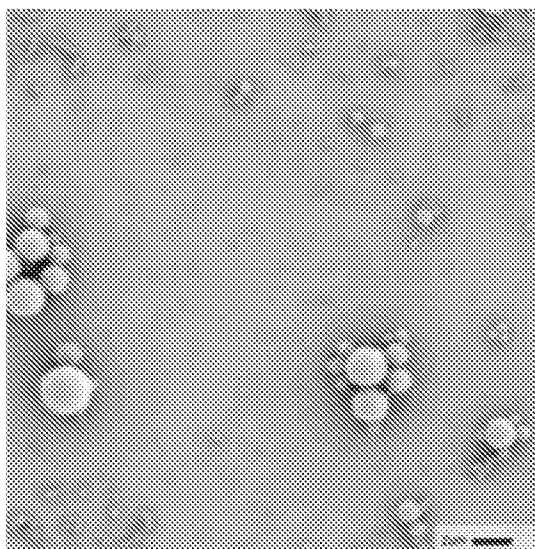


FIG. 13A

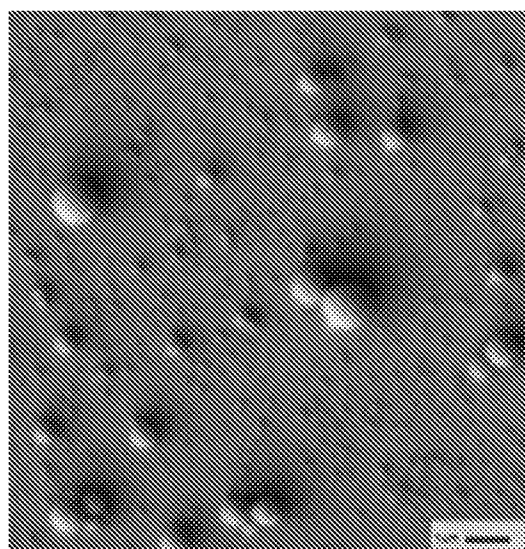


FIG. 13B

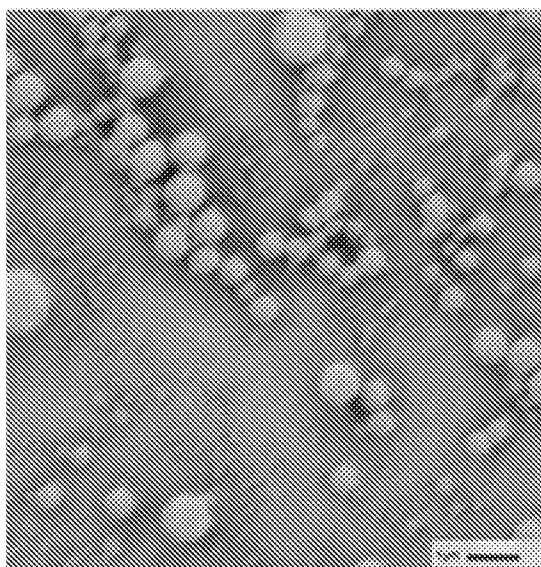


FIG. 13C

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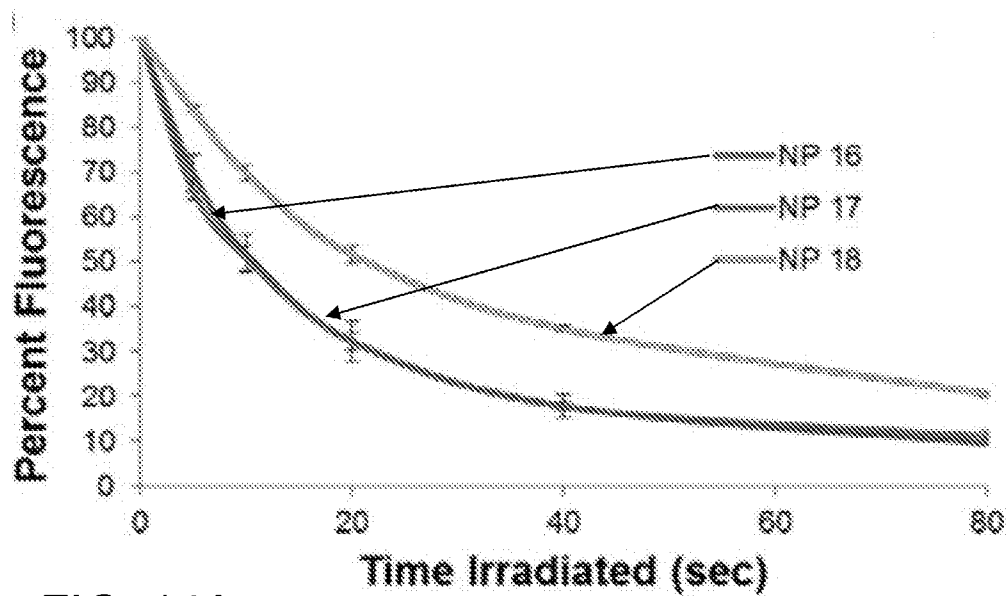


FIG. 14A

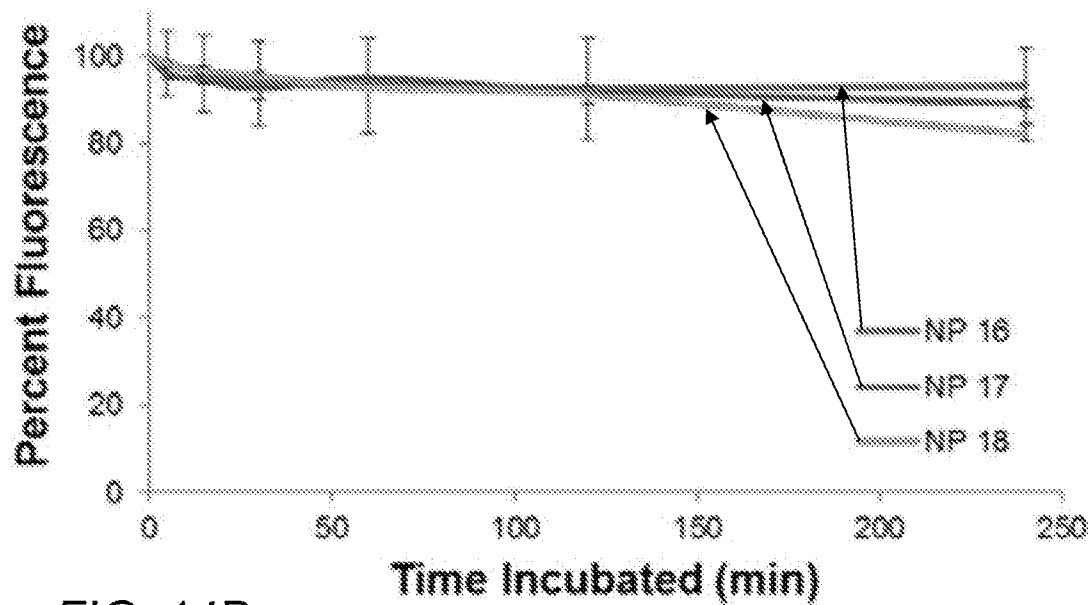


FIG. 14B

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US2016/026792

A. CLASSIFICATION OF SUBJECT MATTER

IPC(8) - C07D 307/33; A61K 9/16; C07C 29/60, C07C 29/132; C07G 3/00 (2016.01)

CPC - C07D 307/33; A61K 9/1647; C07C 29/60, C07C 29/132; C07G 3/00 (2016.05)

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC(8) - A61K 9/16; C07C 29/60, C07C 29/132; C07D 307/33; C07G 3/00 (2016.01)

CPC - A61K 9/1647; C07C 29/60, C07C 29/132; C07D 307/33; C07G 3/00 (2016.05)

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

USPC: 514/175, 514/772.3; IPC(8) - A61K 9/16; C07C 29/60, C07C 29/132; C07D 307/33; C07G 3/00 (2016.01); CPC - A61K 9/1647; C07C 29/60, C07C 29/132; C07D 307/33; C07G 3/00 (2016.05) (keyword delimited)

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

Orbit, Google Patents, Google Scholar, Google, PubChem

Search terms used: poly(lactide-co-glycolide), PLGA, deprotect, light, depolymerize, intramolecular cyclization

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	LUX et al.: Intramolecular Cyclization for Stimuli-Controlled Depolymerization of Polycaprolactone Particles Leading to Disassembly and Payload Release. ACS Macro Lett.; 2 (5): 432-435, 2013. [retrieved on 02.06.2016]. Retrieved from the Internet. <URL: http://pubs.acs.org/doi/abs/10.1021/mz400129h >. entire document	1-2b
A	MAHATO et al.: Biomaterials for Delivery and Targeting of Proteins and Nucleic Acids. CRC Press, 2005. [retrieved on 02.06.2016]. Retrieved from the Internet. <URL: http://dlia.ir/Scientific/e_book/Medicine/Pharmacy_and_Materia_Medica/RS_200_201_Pharmaceutical_Dosage_Forms_/019484.pdf >. Sections 3.2.2 3.5, 4.1-4.4.2	1-25
A	US 2013/0236500 A1 (BIND BIOSCIENCES INC) 12 September 2013 (12.09.2013) entire document	1-25
A	WO 2011/156686 A2 (THE REGENTS OF THE UNIVERSITY OF COLORADO, A BODY CORPORATE) 15 December 2011 (15.12.2011) entire document	1-25
P, X	OLEJNICZAK et al.: LightTriggered Intramolecular Cyclization in Poly(lacticoglycolic acid) Based Polymers for Controlled Degradation. Macromolecules, 48 (10), pp 3166-3172, 2015. [retrieved on 02.06.2016]. Retrieved from the Internet. <URL: http://pubs.acs.org/doi/abs/10.1021/acs.macromol.5b00455 >. Abstract and Supporting Information	1-25

☐ Further documents are listed in the continuation of Box C.☐ See patent family annex.

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"E" earlier application or patent but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

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Date of the actual completion of the international search

02 June 2016

Date of mailing of the international search report

05 JUL 2016

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