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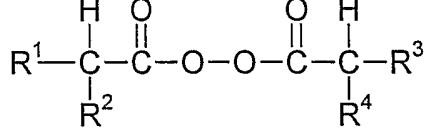
(54) Title: POLYMERIZATION PROCESS INVOLVING DIACYL PEROXIDES

(57) Abstract: The invention relates to a process to make and use specific diacyl peroxides in polymerization reactions. The peroxides are produced and used in a polymerization process within a short timeframe in order to overcome stability issues associated with them.

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POLYMERIZATION PROCESS INVOLVING DIACYL PEROXIDES

The present invention relates to a polymerization process wherein at least one

diacyl peroxide of formula (I)  is used as a source of

5 free radicals.

Such processes are known from, for instance, WO 00/17245, where diisobutanoloyl peroxide is used to polymerize vinyl chloride by means of the radicals that are formed during the thermal decomposition of the peroxide.

10

However, the use of diacyl peroxides of formula (I) can be cumbersome. More particularly, the peroxides of formula (I) are very thermally labile, as a result of which they show significant decomposition during storage at temperatures as low as -5°C. In order to ensure the quality and assay of the product, it is typically produced at temperatures of 0°C or lower and kept at temperatures of -20°C during its storage and handling. However, even at such low temperatures, a noticeable reduction of the assay of the material is noted. Furthermore, it is necessary to phlegmatize the peroxides of formula (I), particularly the products wherein R¹, R², R³, and R⁴ contain less than 20 carbon atoms in total, using a water-immiscible, hydrophobic solvent, such as isododecane, because the peroxides cannot be handled in the pure form for safety reasons. Such solvents will end up in the polymer made in the polymerization process, which for a variety of well-known reasons (undesired plasticizing of the polymer, exudation from finished polymer articles, which may lead to fogging, and the like), is typically 15 undesired.

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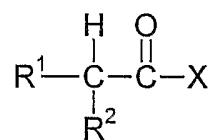
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Therefore, there is need for a process wherein peroxides of formula (I) can be used without undesired refrigeration and solvents being obligatory. Preferably, the process also allows the use of water to reduce the safety hazards associated with the diacyl peroxides, so that they can be handled safely during processing and 5 handling/metering to the polymerization reactor.

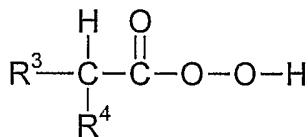
Surprisingly, we have found that the use of a specific process for making and using the diacyl peroxides of formula (I) fulfills this need.

10 The process according to the invention is characterized in that

- an aqueous mixture comprising a peroxide of formula (I) is produced in a peroxidation step wherein



i) one or more acid halides of formula (II) are reacted with
 ii) MOOH/M₂O₂ and/or one or more peracids (or their salts) of formula (III)



15 , in an aqueous phase,

in which peroxidation step the acid halide, or mixture of acid halides, is essentially only brought into contact with water containing MOOH/M₂O₂ and/or one or more peracids (or their M salts), to give an aqueous mixture, and

iii) optionally one or more colloids and/or surfactants are combined with said

20 aqueous mixture before, during, or after the peroxidation step and

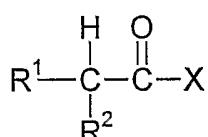
- the resulting aqueous mixture is used in a polymerization process.

It is noted that the term MOOH/M₂O₂ stands for the product that is formed from H₂O₂ and a suitable source of metal (M) ions. The product typically is not a 25 discrete product M₂O₂, but the equilibrium comprising H₂O₂, MOOH and M₂O₂.

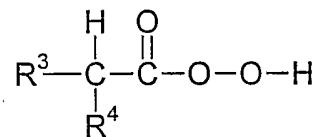
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From an industrial point of view it is preferred that the starting material from which the diacyl peroxides of formula (I) are prepared, i.e. the one or more acid halides



of formula (II) R^2 and the one or more peracids (or their M salts) of



5 formula (III) R^1 , are easily available or obtainable. Most preferably, these compounds are commercially available.

The process of this invention is pre-eminently suitable for diacyl peroxides that are liquid at the process temperature, since these diacyl peroxides have a lower hydrolytic and thermal stability than diacyl peroxides that are solid at the process temperature

In a preferred embodiment, the aqueous mixture comprising diacyl peroxide of formula (I) is used in the polymerization process within a period of 168, more preferably 102, even more preferably 48, more preferably still 24, yet more preferably 12, yet more preferably still 2 hours after the peroxidation step. In a most preferred embodiment, the reactor wherein the peroxidation step is conducted is directly linked by means of piping and optional further holding or processing tanks to the reactor in which the polymerization process is conducted.

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In another preferred embodiment, the diacyl peroxide is stored and handled in this process at temperatures up to 5°C, since it was observed that, depending on the hydrolytic and thermal stability of the diacyl peroxide, it can be acceptable in the process according to the invention to store and handle the diacyl peroxide up to 25 temperatures of -5°C, preferably 0°C or more preferably 5°C.

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In the process of the present invention it is necessary to prevent just water and acid halide of formula (II) from being combined at any given time, since otherwise the acid halide would hydrolyze, resulting in low yields of peroxide and causing

5 contamination of the aqueous mixture. Also, a mixture of just water and acid halide was found to be very corrosive, which would require the use of more expensive process equipment. Hence, when it is stated that in the process the acid halide is essentially only brought into contact with water containing MOOH/M₂O₂ and/or one or more peracids, what is meant is that water and just acid halide are contacted for

10 a period of at most 1 minute, preferably at most 20 seconds, more preferably at most 10 seconds, even more preferably at most 5 seconds, most preferably at most 1 second. Therefore, the only viable ways to make diacyl peroxide are: i) to pre-charge water, MOOH/M₂O₂ (or, optionally, to make the MOOH/M₂O₂ from H₂O₂ and a source of M) and dose acid halide thereto, ii) to pre-charge water and H₂O₂

15 and dose both the source of M and acid halide, iii) to pre-charge water and dose both MOOH/M₂O₂ (optionally H₂O₂ together with a source of M) and acid halide, and iv) to pre-charge water and a source of M and dose both H₂O₂ and acid halide. If a peracid is used in the present process, the same viable ways are available (substituting the peracid for H₂O₂, and substituting the M salt of the peracid for

20 MOOH/M₂O₂). In order to optimize the yield of peroxide, using one of the four viable ways is preferred.

It is noted that US 3,923,766 discloses a so-called "*in-situ*" process where a diacyl peroxide of formula (I) is used that is produced from an anhydride and a peracid.

25 The "*in-situ*" process is characterized in that the diacyl peroxide is produced in the polymerization reactor in the presence of the monomer, which is not a process according to the invention. Such "*in-situ*" processes are undesired, since they do not allow a flexible mode of operation. More particularly, most of the peroxide is formed near the start of the process, preventing a constant polymerization rate,

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and thus resulting in a heat output that is not constant. Such a variation in heat output is undesired, since polymerization reactors are run most cost-effectively under conditions where cooling is at the maximum rate. To overcome this problem, it is suggested in US 3,923,766 to use a second (conventional) more stable

5 initiator. However, such more stable initiators will partially end up in the final polymer, which is undesired for polymer stability reasons (the remaining initiator will lead to premature decomposition of the resin during heat treatment, such as moulding operations).

10 US 3,936,506 discloses a method of preparing asymmetrical α -halogen substituted diacyl peroxides of formula (I) wherein R^2 is chloride or bromide, R^4 is hydrogen, and each of R^1 and R^3 is a long chain alkyl radical having about 10 to about 16 carbon atoms. It is noted that these α -halogen substituted diacyl peroxides of US 3,936,506 are not of interest for the present invention, since the starting materials 15 from which these diacyl peroxides are to be prepared are not easily available or obtainable. From an industrial point of view this is an undesirable situation.

It is furthermore noted that WO 01/32613 discloses a process where peroxydicarbonates are produced *ex-situ*. However, it is desired to offer alternative 20 processes allowing the production of a broader range of peroxides. More specifically, the peroxydicarbonates of WO 01/32613 all have a half-life of 1 hour at a temperature of about 64°C. This means a rather limited freedom when it comes to selecting the polymerization temperature. Also, the chloroformate used to make the peroxydicarbonates is much more hydrolytically stable than the acid 25 halides that are used to make the diacyl peroxides of the present invention. At lower temperatures (say about 50°C) the peroxydicarbonates will not decompose fast enough to give efficient polymerization times and lead to an unacceptably high residual peroxide content in the final polymer, which adversely affects the thermal stability of the formed polymer. At higher polymerization temperatures (say about

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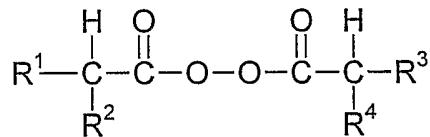
60°C) the peroxydicarbonates will decompose too quickly, leading to inefficient use and too slow polymerization rates towards the end of the polymerization. Surprisingly, it was found that, after further modification, the concept of WO 01/32613 could be extended to diacyl peroxides of formula (I). The modification

5 needed is, *inter alia*, to make sure that the acid halide is not just substituted for the chloroformate, but that it is brought into contact with water only when an inorganic peroxide or peracid (salt) is present in the aqueous phase.

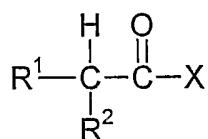
Detailed description of the invention

10 The present invention relates to a process comprising the steps of:

- producing an aqueous mixture comprising a diacyl peroxide of formula (I)



, wherein R¹-R⁴ are independently selected from the group consisting of hydrogen, halogen, and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or 15 more halogen atoms, provided that if R¹ and/or R² are hydrogen, then R³ and/or R⁴ are not hydrogen and, vice-versa, that if R³ and/or R⁴ are hydrogen, then R¹ and/or R² are not hydrogen, in which process R¹-C-R² and/or R³-C-R⁴ can be part of a ring structure that can be saturated or unsaturated and substituted with one or more independently chosen groups R¹, or one of R¹-C(R²)H and R³-C(R⁴)H represents an, optionally substituted, aromatic ring 20 structure, by reacting, in a so-called peroxidation step, one or more acid halides

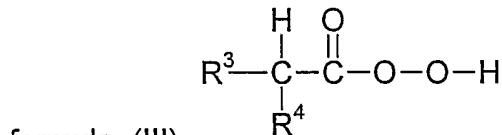


of formula (II) , wherein X is halide, preferably chloride or bromide, most preferably chloride, with i) MOOH/M₂O₂, wherein M is any metal or ammonium-containing group that will react with H₂O₂ to form MOOH/M₂O₂ 25 without decomposing one or more of the peroxides present in the process,

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preferably M is selected from the group consisting of ammonium, sodium, potassium, magnesium, calcium, and lithium, and/or ii) one or more peracids of



, or preferably the M salts thereof, in an aqueous phase, in which process the acid halide, or mixture of acid halides, is essentially only brought into contact with water containing MOOH/M₂O₂ and/or one or more peracids (or their peracid salts), preferably M₂O₂ or the M salt of the peracid,

- 5 - optionally introducing one or more solvents for the acid halide in any part of the process,
- 10 - optionally introducing one or more salts in any part of the process,
- optionally introducing one or more colloids and/or surfactants, before, during, or after the peroxidation step,
- optionally purifying the aqueous mixture in one or more purification steps,
- optionally homogenizing the aqueous mixture in one or more homogenization
- 15 steps,
- transferring the product from the previous steps that comprises the diacyl peroxide of formula (I) to a polymerization reactor, and
- thermally decomposing said diacyl peroxide, to generate organic free radicals, in the presence of one or more ethylenically unsaturated monomers, to
- 20 polymerize said monomers in said polymerization reactor.

In a preferred embodiment, the aqueous mixture comprising said diacyl peroxide is used in said polymerization process within a period of 168, more preferably 102, even more preferably 48, more preferably still 24, yet more preferably 12, yet more preferably still 2 hours after said peroxidation step. In a most preferred embodiment, the reactor wherein the peroxidation step is conducted is directly linked by means of piping and optional further holding or processing tanks to the

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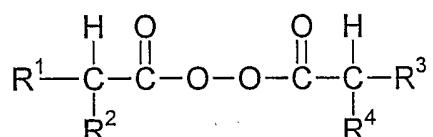
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reactor in which the polymerization process is conducted. Preferably, the diacyl peroxide-containing mixture is prepared just prior to when it is needed in the polymerization step. Should the dispersion be stored for any period of time, then it is preferred, from a safety point of view, to keep the dispersion in motion to prevent phase separation. Any conventional agitation can be used, such as a shaft with blades, a system to bubble inert gas through the dispersion, and/or a recirculation pump. If one or more of the optional colloids and/or surfactants are used in the process, then the diacyl peroxide-containing mixture typically is less prone to phase separation, increasing safety margins.

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In another preferred embodiment, the diacyl peroxide is stored and handled in this process at temperatures up to 5°C, since it was observed that, depending on the hydrolytic and thermal stability of the peroxide, it can be acceptable to store and handle the peroxide in the process according to the invention up to temperatures of -5°C, preferably 0°C or more preferably 5°C.

In a further preferred embodiment, the process involves diacyl peroxides of



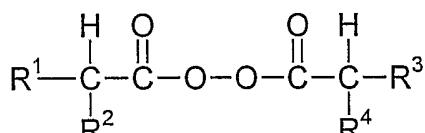
20 formula (I) R^1 R^2 R^3 R^4 , wherein R^1 - R^4 are independently selected from the group consisting of hydrogen, halogen, and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or more halogen atoms, provided that if R^1 and/or R^2 are hydrogen, then R^3 and/or R^4 are not hydrogen and, vice-versa, that if R^3 and/or R^4 are hydrogen, then R^1 and/or R^2 are not hydrogen, in which process R^1 - C - R^2 and/or R^3 - C - R^4 can be part of a ring structure that can be saturated or unsaturated and substituted with one or more independently chosen groups R^1 , or one of R^1 - $C(R^2)H$ and R^3 - $C(R^4)H$ represents an, optionally substituted, aromatic ring structure, with the further proviso that diacyl peroxides of formula (I) wherein R^2 is

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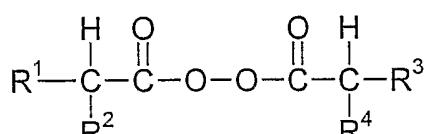
halogen and R⁴ is hydrogen, and both R¹ and R³ are long chain alkyl groups having about 10 to about 16 carbon atoms, are excluded from this invention.

In yet another preferred embodiment, the process involves diacyl peroxides of



5 formula (I) R^2 R^4 , wherein R^1 - R^4 are independently selected from the group consisting of hydrogen, halogen, and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or more halogen atoms, provided that if R^1 and/or R^2 are hydrogen, then R^3 and/or R^4 are not hydrogen and, vice-versa, that if R^3 and/or R^4 are hydrogen, then R^1 and/or R^2 are not hydrogen, in which process R^1 - C - R^2 and/or R^3 - C - R^4 can be part of a ring structure that can be saturated or unsaturated and substituted with one or more independently chosen groups R^1 , or one of R^1 - $C(R^2)H$ and R^3 - $C(R^4)H$ represents an, optionally substituted, aromatic ring structure, with the further proviso that if R^2 is halogen and R^4 is hydrogen, and both R^1 and R^3 are alkyl groups, then each of these alkyl groups has up to about 6 carbon atoms.

In yet another preferred embodiment, the process involves diacyl peroxides of



20 formula (I) R^2 R^4 , wherein R^1 - R^4 are independently selected from the group consisting of hydrogen and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or more halogen atoms, provided that if R^1 and/or R^2 are hydrogen, then R^3 and/or R^4 are not hydrogen and, vice-versa, that if R^3 and/or R^4 are hydrogen, then R^1 and/or R^2 are not hydrogen; in which process R^1 -C- R^2 and/or R^3 -C- R^4 can be part of a ring structure that can be saturated or unsaturated and substituted with

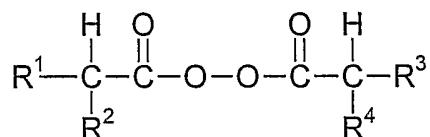
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one or more independently chosen groups R^1 , or one of $R^1\text{-}C(R^2)\text{H}$ and $R^3\text{-}C(R^4)\text{H}$ represents an, optionally substituted, aromatic ring structure.

5 In yet another preferred embodiment of this invention, the diacyl peroxide of formula (I), when transferred to the polymerization reactor, is always transferred from the previous step in the aqueous phase, with the diacyl peroxide being



10 represented by formula (I), wherein $R^1\text{-}R^4$ are independently selected from the group consisting of hydrogen, halogen, and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties 15 optionally substituted with one or more halogen atoms, provided that if R^1 and/or R^2 are hydrogen, then R^3 and/or R^4 are not hydrogen and, vice-versa, that if R^3 and/or R^4 are hydrogen, then R^1 and/or R^2 are not hydrogen, in which process $R^1\text{-}C\text{-}R^2$ and/or $R^3\text{-}C\text{-}R^4$ can be part of a ring structure that can be saturated or unsaturated and substituted with one or more independently chosen groups R^1 , or 15 one of $R^1\text{-}C(R^2)\text{H}$ and $R^3\text{-}C(R^4)\text{H}$ represents an, optionally substituted, aromatic ring structure.

20 In one still more preferred embodiment only one reactor is used to produce the diacyl peroxide. Most preferably, only one reactor is used to react both the source of M and H_2O_2 and to carry out the peroxidation step to make the diacyl peroxide.

25 The amount of diacyl peroxide (which acts as an initiator) to be used as a source of free radicals in the polymerization step according to the invention is within the range conventionally used in polymerization processes. Typically, from 0.005 to 2% by weight (%w/w) of initiator, preferably 0.01-1 %w/w, more preferably 0.01-0.5 %w/w, based on the weight of the ethylenically unsaturated monomer(s) to be polymerized, is used. It is noted that the diacyl peroxide initiator of the present invention may be

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used in combination with other initiators.

Preferably, the acid halide of formula (II) is derived from C₁-C₃₀ carboxylic acids, preferably from isobutanoic acid, 2-methylbutanoic acid, 2-methylpentanoic acid,

5 2-methylhexanoic acid, 2-methylheptanoic acid, octanoic acid, 2-methyloctanoic acid, decanoic acid, 2-methylnonanoic acid, tetradecanoic acid, hexadecanoic acid, octadecanoic acid, lauric acid, 2-ethylbutanoic acid, 2-ethylpentanoic acid, 2-ethylhexanoic acid, 2-ethylheptanoic acid, 2-ethylnonanoic acid, 2-propylcaprylic acid, 2-propylheptanoic acid, 2-propylhexanoic acid, 2-propyoctanoic acid, 2-propylpentanoic acid, 2-propyl-4-pentanoic acid, 2-butyldcaproic acid, 2-butylheptanoic acid, 2-butyloctanoic acid, (substituted) cyclohexane carboxylic acid, chloroacetic acid, 2-chloropropionic acid, 2-chlorobutanoic acid, 2-chlorohexanoic acid, dichloroacetic acid, propanoyl-C₁₆ acid, and aromatic acids such as benzoic acid, 2-chlorobenzoic acid, 3-chlorobenzoic acid, 2-methylbenzoic acid, 3-methylbenzoic acid, 4-methylbenzoic acid, naphtoic acid, and toluic acid. More preferably, at least one of the acid halides is selected from isobutanoyl halide, 2-ethylbutanoyl halide, 2-ethylhexanoyl halide, 2-chloropropanoyl halide, dichloroacetic acid halide, lauroyl halide, cyclohexane carbonyl halide, 3-methylcyclohexane carbonyl halide, 2,3 dihalocyclohexane carbonyl halide, 20 benzoyl halide, 2-methylbenzoyl halide, and 2-chlorobenzoyl halide. If so desired, also the acid halides of polycarboxylic acids, with two or more acid moieties, can be used, such as (cyclo) hexyl dicarboxylic acid, phthalic acid (any isomer), maleic acid, fumaric acid, 1,2,4-butanetricarboxylic acid, (oligomeric) polyacrylic acid, and copolymers with maleic anhydride derivable acid groups. Suitably, a mixture of any 25 of these acid halides is used, typically resulting in a mixture of three or more diacyl peroxides of formula (I), as is known in the art.

If the acid halide of formula (II) is reacted with a peracid of formula (III), or the salt of such a peracid, wherein R³ and R⁴ are not hydrogen, then also acid halides of formula (II) can be used wherein R¹ and/or R² are hydrogen. Examples of such

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acid halides include chloroacetic acid halide, chloropropionic acid halides (all isomers), acetyl chloride, propionyl chloride, butanoyl chloride, pentanoyl or valeroyl chloride, decanoyl chloride, lauroyl chloride, malonyl chloride, succinoyl chloride, glutaroyl chloride, adipoyl chloride, azeloyl chloride, sebacoyl chloride.

5 Most preferably, the diacyl peroxide produced is diisobutanoyl peroxide.

The M_2O_2 is preferably prepared by combining hydrogen peroxide and NaOH, KOH, K_2CO_3 , and/or Na_2CO_3 . The M_2O_2 can be prepared separately or in a reaction step of the present process.

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Any peracid used can be prepared in a conventional way. Preferably, a C_2-C_{20} peracid is used. It is noted that when a peracid is reacted with acid halide, typically an asymmetrical diacyl peroxide is obtained, meaning that R^1 and R^2 are not the same as R^3 and R^4 , whereas if a mixture of two different acid halides is reacted with $MOOH/M_2O_2$, a statistical mixture of two symmetrical and one asymmetrical diacyl peroxides is obtained. The use of a peracid allows better control of the polymerization rate, since just one product with a dedicated half-life time is formed. Furthermore, it is noted that if asymmetrical diacyl peroxide is desired of which R^3 or R^4 is hydrogen, the use of a mixture of acid halides will partially result in the formation of diacyl peroxide that shows no branching on the two α -carbons (relative to the diacyl peroxide function). Such α -carbon unbranched diacyl peroxides are quite stable and part of them will end up in the final polymer, typically resulting in a lower than desired heat stability of the polymer. Hence, it is preferred to use a peracid if it is desired that R^3 and/or R^4 are hydrogen. More preferably, the peracid is selected from the group consisting of peracetic acid, perpropionic acid, perisobutanoic acid, perhexanoic acid, perbenzoic acid, 3-chloroperbenzoic acid, and perlauric acid, the use of perpropionic acid and perlauric acid being most preferred.

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Preferably, the process is conducted such that it is essentially free of solvent, since such solvents are not desired in the final polymer. For the purpose of this specification, essentially solvent-free means that less than 20 %w/w of solvent, based on the weight of the diacyl peroxide, is present. Preferably less than 10
5 %w/w, more preferably less than 5 %w/w, most preferably less than 2 %w/w of solvent is present, all based on the weight of the diacyl peroxide. However, if the use of certain solvents, such as conventional plasticizers for PVC or phlegmatizers for the peroxide, is acceptable in the polymer obtained in the polymerization step, it may be advantageous to use such solvents. Conventional plasticizers include
10 epoxidized soybean oil, dialkyl esters, such as alkyl esters of aliphatic carboxylic acids with two or more carboxylic acid moieties, and less desired phthalate esters. Conventional phlegmatizers include hydrocarbons, such as isododecane.

Preferably, the peroxidation step of the process is conducted without a salt being
15 added to the reaction mixture. However, if so desired, a salt may be added, for instance to increase the yield of diacyl peroxide. If used, the salt is preferably selected from alkali and/or alkaline-earth metals. Preferably, they are salts of strong acids that do not react with H₂O₂, peracid, or a salt thereof. More preferably, the salt is selected from NaCl, KCl, Na₂SO₄, K₂SO₄, and NH₄Cl.

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If used, the colloids are advantageously selected from the group consisting of hydrolyzed polyvinyl acetate, alkyl cellulose, hydroxyalkyl alkyl cellulose, gelatin, polyvinyl pyrrolidone, polyoxyethylene sorbitan monolaurate, and polyacrylic acid. Preferably, said dispersant is a non-ionic compound. Most preferably, it is a
25 mixture of one or more hydrolyzed polyvinyl acetates (PVA). Suitably, the degree of hydrolysis of the PVA ranges from 50 up to 95%, preferably up to 90%. Preferably, the degree of hydrolysis is at least 55%, more preferably at least 60%, so that the PVA is soluble in the diacyl peroxide-containing mixture. In order to

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increase the solubility of the PVA, it can be advantageous to add a C₁-C₄ alcohol, preferably methanol, ethanol, and/or isopropanol.

All conventional surfactants can be used. However, in order not to disturb the dispersion of the polymerization reaction and not to adversely affect the properties of the final polymer, the surfactant preferably is a (biodegradable) cationic compound, such as preferred quaternary ammonium compounds, or a non-ionic surfactant. Preferred conventional non-ionic surfactants include, but are not limited to, alkoxylated (fatty) alcohols and alkoxylated sugars.

10

The aqueous mixture that results from the peroxidation step can be transferred to the polymerization reactor "as is". However, in order to enhance the safety of handling of the diacyl peroxide, it is preferred to add at least one surfactant and/or colloid, so that a physically stable dispersion of the diacyl peroxide in the aqueous mixture, or a dispersion of the aqueous phase in the diacyl peroxide, is obtained. If need be, the size of the dispersed phase (droplet or particle size) can be reduced in an optional further homogenization step. However, such a step is only a prerequisite when the dispersion as obtained is not safe. Suitably, this is detected by heating a 100 g sample of the dispersion up to a temperature where the diacyl peroxide decomposes. If a certain phase separation is observed, the safety characteristics of the least safe phase are decisive. However, it may also be desired to obtain a very finely dispersed phase, for instance to enhance an even distribution of the diacyl peroxide over the monomer(s) in the polymerization reactor (e.g. to reduce the fish-eye level in PVC that is produced by a suspension polymerization process). If such fine dispersions (with the average size by volume of the dispersed phase being less than 10 µm, preferably less than 5 µm, most preferably less than 2 µm) are of interest, then also an extra homogenization step can be advantageously introduced. The homogenization step can be performed

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using any suitable (high) shear mixer, such as rotor/stator homogenizers, colloid mills, ultrasonic devices, and the like.

In a preferred embodiment, the diacyl peroxides of this invention are prepared in a
5 feed line, such as a pipe, that is directly connected to the polymerization reactor. The desired (raw) materials for the peroxidation step are then fed into the pipe at suitable points. Optionally, an in-line mixer is then used to homogenize the resulting product of the peroxidation step before it is used in the polymerization step. If so desired, part of the product of the peroxidation step may be recycled in
10 such a set-up to increase the yield of diacyl peroxide. If so desired, the line may be flushed clean with water after the peroxidation step. However, it can also be switched to another polymerization reactor, such that the peroxidation can be an essentially continuous operation.

15 The product resulting from the peroxidation step typically contains by-products and residual raw materials. Particularly when the diacyl peroxide is prepared using MOOH/M₂O₂, a substantial quantity of a salt (MX) is formed. The by-products must not be a hindrance in the polymerization reaction. However, it is known that salts
20 adversely affect the stability of the emulsion of the diacyl peroxide, which is a safety concern, while salts may also adversely affect the electrical properties of the final polymer. Hence, it may be preferred to include a purification step. Depending on safety considerations, it can be acceptable to conduct the peroxidation step without any surfactants and/or colloids being used and then to separate and discard (part of) the water layer, after which the diacyl peroxide-rich organic layer
25 may be dispersed and/or diluted, if so desired, using the appropriate media and/or solvents. If an aqueous diacyl peroxide dispersion is to be achieved, conventional surfactants and/or colloids can be used in the dispersion step. Alternatively, (part of) the salt may be removed using conventional techniques, such as (reverse) osmosis, nanofiltration, ion-exchange, precipitation, and the like. A less preferred

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process would involve a step wherein a conventional solvent for the diacyl peroxide is added to extract the diacyl peroxide from the mixture.

The process according to the invention is pre-eminently suited to polymerize one or 5 more ethylenically unsaturated monomers, including (meth)acrylic acid (esters), styrene, vinyl acetate, acrylonitrile, vinyl chloride monomer (VCM), and the like. Preferably, the process according to the invention involves the polymerization of monomer mixtures comprising at least 50 %w/w of VCM, based on the weight of all monomer(s). In this preferred process, preferred comonomers for use are of the 10 conventional type and include vinylidene chloride, vinyl acetate, ethylene, propylene, acrylonitrile, styrene, and (meth)acrylates. More preferably, at least 80 %w/w of the monomer(s) being polymerized is made up of VCM, while in the most preferred process the monomer consists essentially of VCM. As is known in the art, the polymerization temperature of such processes determines to a large extent the 15 molecular weight of the final polymer.

The polymerization process can be conducted as a mass process where the reaction mixture is predominantly monomer(s) or as a suspension process where the reaction mixture typically is a suspension of monomer(s) in water, or as an emulsion 20 or micro-emulsion process where the monomer(s) typically is/are emulsified in water. In these processes the usual additives will have to be used. For example, if the monomer(s) is/are present in the form of a suspension in water, the usual additives such as surfactant(s), protective colloid(s), anti-fouling agent(s), pH-buffer(s), etc. may be present. Depending on the type of polymer desired, each of the above- 25 mentioned processes may be preferred. The process according to the invention is especially suited for mass and suspension processes. In the aqueous suspension process to produce PVC from VCM, the polymerization is usually conducted at a temperature in the range of about 0°C to 100°C. However, for the process of the present invention it is preferred to employ polymerization temperatures of about

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40°C to about 70°C, since this is the temperature at which VCM is polymerized efficiently. The polymerization reaction time may vary from about 1 to about 15 hours, and is preferably from 2 to 6 hours. The aqueous suspension VCM polymerization process in addition to the VCM typically contains water, dispersants,

5 free radical initiator, and optional further ingredients, such as buffers, short stop agents, pre-stabilizers, and the like.

After the polymerization, the resulting (co)polymer (or resin) will be worked up as is usual in the art. Polymers obtained by a suspension polymerization according to the

10 invention, for example, will be submitted to the usual drying and screening steps.

The resulting resin preferably is characterized in that it contains less than 50 ppm of residual diacyl peroxide, more preferably less than 40 ppm, and most preferably less than 25 ppm of diacyl peroxide, immediately after drying for 1 hour at 60°C and screening. The resin was found to exhibit excellent heat stability, as measured with a

15 Metrastat® PSD260 testing oven in accordance with ISO 182-2 (1990E). The improved heat stability proved that the resin hardly discoloured when submitted to melt-processing steps, e.g., to form shaped articles.

The process according to the present invention is pre-eminently suited for

20 combination with polymerization processes as disclosed in WO00/17245.

Accordingly, a preferred embodiment of the present invention relates to a process of claim 1 wherein at least part of the diacyl peroxide is dosed to the polymerization mixture at the reaction temperature. More preferably, essentially all of the organic initiator, i.e. diacyl peroxides and, if used, other conventional initiators, dosed in the

25 polymerization process has a half-life of from 0.0001 hour to 1.0 hour at the polymerization temperature, preferably from 0.001 to 0.8 hour, more preferably from 0.002 to 0.5 hour at the polymerization temperature. If the one or more diacyl peroxides are dosed together with other initiators, it is preferred that the diacyl peroxide(s) as well as all other initiators fulfill these half-life requirements. It is noted

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that half-lives are determined in conventional thermal decomposition studies in monochlorobenzene, as is well-known in the art (see for instance the brochure "Initiators for high polymers" with code 10737 obtainable from Akzo Nobel).

- 5 The dosing of the extremely fast initiators to the polymerization reaction mixture can be intermittent or continuous over a period of time wherein at least 20%, preferably at least 40%, more preferably at least 60% of all monomer used in the process is polymerized. If an intermittent operation is selected, there are at least 2, preferably at least 4, more preferably at least 10, and most preferably at least 20 moments at
- 10 the polymerization temperature at which the initiator is dosed. If so desired, the intermittent and continuous operations may be combined, such that the initiator is dosed intermittently for certain (longer or shorter) periods of time. Most preferably, the diacyl peroxide is dosed continuously and/or intermittently after at least 5%, preferably at least 10%, more preferably at least 20%, most preferably at least 30%
- 15 of the monomer(s) has already been polymerized and while during the dosing period at least 10%, preferably at least 20%, more preferably at least 30%, and most preferably at least 50% of all monomer used in the process is polymerized. Preferably, the diacyl peroxide is dosed in a concentration of 0.1 to 60 %w/w, more preferably 0.5 to 25 %w/w, and most preferably 2 to 15 %w/w, based on the weight
- 20 of the diacyl peroxide-containing mixture, in order to facilitate the even distribution of the diacyl peroxide over the monomer.

- 25 The dosing can be effected at any suitable entry point to the reactor. If water is added in the course of the polymerization process, for example to compensate for the shrinkage of the reactor content due to the polymerization reaction, it may be advantageous to use the line through which this water is dosed to also dose the initiator. In a most preferred embodiment, the diacyl peroxide has also been produced in said line.

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Experimental

The following chemicals were used in the process to manufacture diacyl peroxides. It is noted that all experiments were performed taking sufficient safety measures. In this respect it is also noted that pure diisobutanoyl peroxide is a detonatable

5 explosive.

Commercially available aqueous H₂O₂ with an assay of 69.97% was used.

NaOH solution (NaOH-25 and NaOH-33) containing 25 %w/w and 33 %w/w, respectively, of NaOH was produced from Baker-grade NaOH and distilled water.

NaCl solution (NaCl-15 and NaCl-25) containing 15 %w/w and 25 %w/w,

10 respectively, of NaCl was produced from Baker-grade NaCl and distilled water.

Na₂CO₃ solution containing 10 %w/w of Na₂CO₃ was produced from Baker-grade Na₂CO₃ and distilled water.

HCl solution containing 18 %w/w of HCl was produced from Baker-grade HCl and distilled water.

15 Isobutanoyl chloride (99.3%) was supplied by BASF

Lauroyl chloride (98%) was supplied by Acros.

A fresh PVA solution, 5 %w/w Alcotex® 72.5 ex Harco in demineralized water was prepared fresh, 1 day prior to use.

20 Example 1

For the preparation of isobutanoyl-lauroyl peroxide, a reactor was charged at 0°C with 35 g deionized water, 286 g NaCl-25 solution, 110 g of Na₂CO₃ solution, 39 g NaOH-25 solution, and 32.5 g H₂O₂ solution. Then, within 50 minutes 103.0 g isobutanoyl chloride were dosed while maintaining a pH-value of 11.0 with the aid 25 of NaOH-25 solution. The temperature was maintained at 0°C, and meanwhile the contents of the reactor were homogenised. After addition of the total amount of isobutanoyl chloride, the reaction mixture was stirred for another 15 minutes at pH 11.0. Then, the pH was lowered to 10.0 with HCl solution and the temperature was brought to 5°C. Next, 70.5 g lauroyl chloride were dosed over a period of 10

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minutes, while maintaining the pH on 10.0 with the aid of NaOH-25 solution. After addition of the total amount of lauroyl chloride, stirring was continued for another 15 minutes.

5 The reaction mixture was set aside for 30 minutes in order to allow the two-phase system to separate. The aqueous layer was removed (713 g) and the organic peroxide layer was washed with 300 g NaCl-15 solution. Again, the aqueous layer and the organic layer were allowed to separate, in order to facilitate the isolation of the peroxide layer (Isolated yield: 145.0 g).

10 Analysis of the isolated peroxide layer:

35% diisobutanyl peroxide
64% isobutanoyl-lauroyl peroxide
<1% dilauroyl peroxide
<0.01% acid chloride

15

Example 2 and Comparative Examples A-D

In Example A, the procedure of WO 01/32613 was repeated. First, a solution of Na₂O₂ was prepared by combining 108.3 g of demineralized water, 18.5 g of NaOH-33 solution, and 3.7 g of the aqueous H₂O₂. A second reactor was charged 20 with 33.3 g of demineralized water, 27.7 g of the PVA solution, and 16.3 g of isobutanoyl chloride while controlling the temperature between 5-10°C. Next, the Na₂O₂ solution was dosed while the system was homogenized. No diisobutanyl peroxide was obtained and only an aqueous phase formed.

In Example B, the same procedure was followed as in Example A, except that 2-ethylbutyroyl chloride was used instead of the isobutanoyl chloride. Again, no diacyl peroxide was obtained.

In Example C, the same procedure was followed as in Example A, except that 2-ethylhexanoyl chloride was used instead of the isobutanoyl chloride. Again, no diacyl peroxide was obtained. The resulting mixture was milky white, probably

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because a lot of 2-ethylhexanoic acid was formed which did not all dissolve in the water.

In Example D, the procedure of Example A was modified in that the second reactor was charged with all the water, PVA solution, isobutyroyl chloride, and aqueous

5 H_2O_2 , and the NaOH-33 solution was then dosed to it at a temperature of 5-10°C.

The yield of diisobutanyl peroxide is expected to be less than 20 %w/w.

In Example 2, a single reactor was charged with 141.5 g of demineralized water, 18.5 g of NaOH-33 solution, 27.7 g of PVA solution, and 3.7 g of the aqueous

10 H_2O_2 , after which 16.3 g of isobutanoyl chloride were dosed to the contents of the

reactor, while maintaining a temperature of 5-10°C and while homogenizing the

contents of the reactor. An emulsion resulted. The yield of diisobutanoyl peroxide

was 82 %w/w. It is expected that if the emulsion is used within 48 hours in a

polymerization reaction, it will show the same polymerization rates for VCM as

were observed when using the same amount of a conventional 30 %w/w

15 diisobutanoyl peroxide solution in isododecane.

Example 3

The product of Example 2 was stored for 1.5 hours at 5°C without being stirred.

The active oxygen content (a measure of the amount of diacyl peroxide of formula

20 (I)) of the emulsion was reduced by 2.5% and a little phase separation occurred. In

Example 3, a reactor was charged with 175 g of demineralized water, 16.4 g of

NaOH-33 solution, 43.8 g of NaCl-25 solution, 250.5 g of Na_2CO_3 solution, and 4.8

g of the aqueous H_2O_2 . Then 23.1 g of isobutanoyl chloride was dosed to the

contents of the reactor, while maintaining a temperature of 5-10°C and

25 homogenizing the contents of the reactor. At the end of the reaction the mixture

was neutralized to pH=7. Taking sufficient safety measures, the two-phase system

was allowed to separate and about 70% by volume of the aqueous phase was

removed. To the organic layer and the remainder of the water layer were added

207.2 g of demineralized water, 109.3 g of PVA solution, 13.7 g of methanol, and

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21.9 g of an aqueous 0.2%w/w Berol® 08 (non-ionic emulsifier ex Akzo Nobel) solution. A dispersion resulted which was stable for more than 2.5 hours and showed good performance in the polymerization process of VCM. The yield of diisobutanol peroxide was 86% on isobutanol chloride and 93% on H₂O₂.

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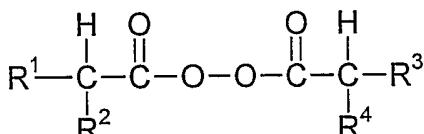
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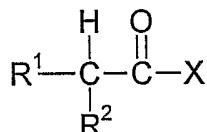
Claims

1. Process comprising the steps of:

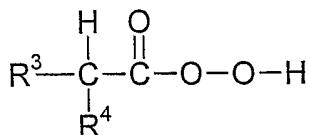
- producing an aqueous mixture comprising a diacyl peroxide of formula (I)



5 , wherein R¹-R⁴ are independently selected from the group consisting of hydrogen, halogen, and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or more halogen atoms, provided that if R¹ and/or R² are hydrogen, then R³ and/or R⁴ are not hydrogen and, vice-versa, that if R³ and/or R⁴ are hydrogen, then R¹ and/or R² are not hydrogen, in which process R¹-C-R² and/or R³-C-R⁴ can be part of a ring structure that can be saturated or unsaturated and substituted with one or more independently chosen groups R¹, or one of R¹-C(R²)H and R³-C(R⁴)H represents an, 10 optionally substituted, aromatic ring structure, by reacting, in a so-called



15 peroxidation step, one or more acid halides of formula (II) , wherein X is halide, preferably chloride or bromide, most preferably chloride, with i) MOOH/M₂O₂, wherein M is any metal or ammonium-containing group that will react with H₂O₂ to form MOOH/M₂O₂ without decomposing one or more of the peroxides present in the process, 20 preferably M is selected from the group consisting of ammonium, sodium, potassium, magnesium, calcium, and lithium, and/or ii) one or more



peracids of formula (III) , or the M salts thereof, in an aqueous phase, in which process the acid halide, or mixture of acid halides,

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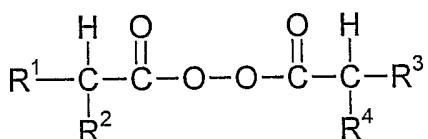
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is essentially only brought into contact with water containing MOOH/M₂O₂ and/or one or more peracids or peracid salts, preferably M₂O₂ or the M salt of the peracid,

- optionally introducing one or more solvents for the acid halide in any part of the process,
- 5 optionally introducing one or more salts in any part of the process,
- optionally introducing one or more colloids and/or surfactants, before, during, or after the peroxidation step,
- optionally purifying the aqueous mixture in one or more purification steps,
- 10 optionally homogenizing the aqueous mixture in one or more homogenization steps,
- transferring the product from the previous steps that comprises the diacyl peroxide of formula (I) to a polymerization reactor, and
- 15 thermally decomposing said diacyl peroxide, to generate organic free radicals, in the presence of one or more ethylenically unsaturated monomers, to polymerize said monomers in said polymerization reactor.

2. Process comprising the steps of:

- producing an aqueous mixture comprising a diacyl peroxide of formula (I)

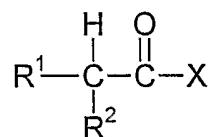


20 , wherein R¹-R⁴ are independently selected from the group consisting of hydrogen, halogen, and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or more halogen atoms, provided that if R¹ and/or R² are hydrogen, then R³ and/or R⁴ are not hydrogen and, vice-versa, that if R³ and/or R⁴ are hydrogen, then R¹ and/or R² are not hydrogen, in which process R¹-C-R² and/or R³-C-R⁴ can be part of a ring structure that can be saturated or unsaturated and substituted with one or more independently

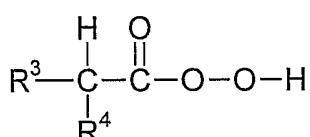
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chosen groups R^1 , or one of $R^1-C(R^2)H$ and $R^3-C(R^4)H$ represents an, optionally substituted, aromatic ring structure, with the further proviso that diacyl peroxides of formula (I) wherein R^2 is halogen and R^4 is hydrogen, and both R^1 and R^3 are long chain alkyl groups having about 10 to about 16 carbon atoms, are excluded, by reacting, in a so-called peroxidation step,



one or more acid halides of formula (II) R^4 , wherein X is halide, preferably chloride or bromide, most preferably chloride, with i) $MOOH/M_2O_2$, wherein M is any metal or ammonium-containing group that will react with H_2O_2 to form $MOOH/M_2O_2$ without decomposing one or more of the peroxides present in the process, preferably M is selected from the group consisting of ammonium, sodium, potassium, magnesium, calcium, and lithium. and/or ii) one or more peracids of formula (III)



R^+ , or the M salts thereof, in an aqueous phase, in which process the acid halide, or mixture of acid halides, is essentially only brought into contact with water containing $MOOH/M_2O_2$ and/or one or more peracids or peracid salts, preferably M_2O_2 or the M salt of the peracid,

- optionally introducing one or more solvents for the acid halide in any part of the process,
- optionally introducing one or more salts in any part of the process,
- 20 - optionally introducing one or more colloids and/or surfactants, before, during, or after the peroxidation step,
- optionally purifying the aqueous mixture in one or more purification steps,
- optionally homogenizing the aqueous mixture in one or more homogenization steps,
- 25 - transferring the product from the previous steps that comprises the diacyl

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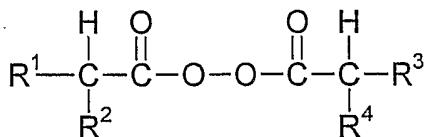
peroxide of formula (I) to a polymerization reactor, and

- thermally decomposing said diacyl peroxide, to generate organic free radicals, in the presence of one or more ethylenically unsaturated monomers, to polymerize said monomers in said polymerization reactor.

5

3. Process comprising the steps of:

- producing an aqueous mixture comprising a diacyl peroxide of formula (I)



, wherein R^1 - R^4 are independently selected

from the group consisting of hydrogen, halogen, and saturated or

10

unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or more halogen atoms, provided that if R^1 and/or R^2 are hydrogen, then R^3 and/or R^4 are not hydrogen and, vice-versa, that if R^3 and/or R^4 are hydrogen, then R^1 and/or R^2 are not hydrogen, in which process R^1 -C- R^2 and/or R^3 -C- R^4 can be part of a ring structure that can be

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saturated or unsaturated and substituted with one or more independently chosen groups R^1 , or one of $R^1-C(R^2)H$ and $R^3-C(R^4)H$ represents an, optionally substituted, aromatic ring structure, with the further proviso that if R^2 is halogen and R^4 is hydrogen, and both R^1 and R^3 are alkyl groups, then each of these alkyl groups has up to about 6 carbon atoms, by reacting, in a so-called peroxidation step, one or more acid halides of formula (II)

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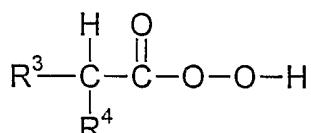
$$\begin{array}{c} \text{H} & \text{O} \\ | & \parallel \\ \text{R}^1 - \text{C} - \text{C} - \text{X} \\ | \\ \text{R}^2 \end{array}$$

, wherein X is halide, preferably chloride or bromide, most
halide, with i) $\text{MOOH/M}_2\text{O}_2$, wherein M is any metal or
containing group that will react with H_2O_2 to form $\text{MOOH/M}_2\text{O}_2$
imposing one or more of the peroxides present in the process,
is selected from the group consisting of ammonium, sodium,

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potassium, magnesium, calcium, and lithium, and/or ii) one or more



peracids of formula (III) R^+ , or the M salts thereof, in an aqueous phase, in which process the acid halide, or mixture of acid halides, is essentially only brought into contact with water containing $\text{MOOH}/\text{M}_2\text{O}_2$ and/or one or more peracids or peracid salts, preferably M_2O_2 or the M salt of the peracid.

10 - optionally introducing one or more solvents for the acid halide in any part of the process,

- optionally introducing one or more salts in any part of the process,

- optionally introducing one or more colloids and/or surfactants, before, during, or after the peroxidation step,

- optionally purifying the aqueous mixture in one or more purification steps,

- optionally homogenizing the aqueous mixture in one or more homogenization steps,

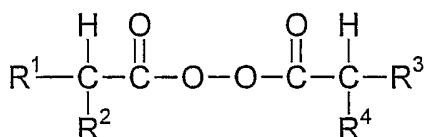
15 - transferring the product from the previous steps that comprises the diacyl peroxide of formula (I) to a polymerization reactor, and

- thermally decomposing said diacyl peroxide, to generate organic free radicals, in the presence of one or more ethylenically unsaturated monomers, to polymerize said monomers in said polymerization reactor.

20

4. Process comprising the steps of:

- producing an aqueous mixture comprising a diacyl peroxide of formula (I)

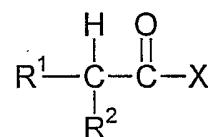


, wherein R'-R' are independently selected from the group consisting of hydrogen and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with

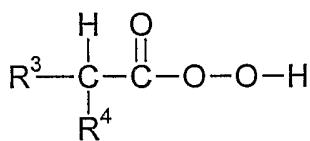
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one or more halogen atoms, provided that if R^1 and/or R^2 are hydrogen, then R^3 and/or R^4 are not hydrogen and, vice-versa, that if R^3 and/or R^4 are hydrogen, then R^1 and/or R^2 are not hydrogen, in which process R^1 -C- R^2 and/or R^3 -C- R^4 can be part of a ring structure that can be saturated or 5 unsaturated and substituted with one or more independently chosen groups R^1 , or one of R^1 -C(R^2)H and R^3 -C(R^4)H represents an, optionally substituted, aromatic ring structure, by reacting, in a so-called peroxidation



step, one or more acid halides of formula (II) , wherein X is halide, preferably chloride or bromide, most preferably chloride, with i) 10 MOOH/ M_2O_2 , wherein M is any metal or ammonium-containing group that will react with H_2O_2 to form MOOH/ M_2O_2 without decomposing one or more of the peroxides present in the process, preferably M is selected from the group consisting of ammonium, sodium, potassium, magnesium, calcium, and lithium, and/or ii) one or more peracids of formula (III)



15 , or the M salts thereof, in an aqueous phase, in which process the acid halide, or mixture of acid halides, is essentially only brought into contact with water containing MOOH/ M_2O_2 and/or one or more peracids or peracid salts, preferably M_2O_2 or the M salt of the peracid,

- 20 optionally introducing one or more solvents for the acid halide in any part of the process,
- optionally introducing one or more salts in any part of the process,
- optionally introducing one or more colloids and/or surfactants, before, during, or after the peroxidation step,
- 25 optionally purifying the aqueous mixture in one or more purification steps,
- optionally homogenizing the aqueous mixture in one or more

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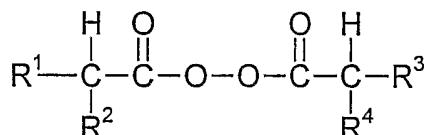
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homogenization steps,

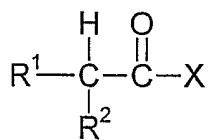
- transferring the product from the previous steps that comprises the diacyl peroxide of formula (I) to a polymerization reactor, and
- thermally decomposing said diacyl peroxide, to generate organic free radicals, in the presence of one or more ethylenically unsaturated monomers, to polymerize said monomers in said polymerization reactor.

5. Process comprising the steps of:

- producing an aqueous mixture comprising a diacyl peroxide of formula (I)



10 , wherein R¹-R⁴ are independently selected from the group consisting of hydrogen, halogen, and saturated or unsaturated, linear or branched alkyl, alkaryl, and aralkyl moieties optionally substituted with one or more halogen atoms, provided that if R¹ and/or R² are hydrogen, then R³ and/or R⁴ are not hydrogen and, vice-versa, that if R³ and/or R⁴ are hydrogen, then R¹ and/or R² are not hydrogen, in which process R¹-C-R² and/or R³-C-R⁴ can be part of a ring structure that can be saturated or unsaturated and substituted with one or more independently chosen groups R¹, or one of R¹-C(R²)H and R³-C(R⁴)H represents an, 15 optionally substituted, aromatic ring structure, by reacting, in a so-called

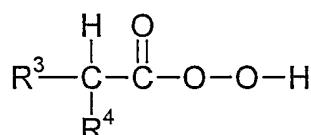


20 peroxidation step, one or more acid halides of formula (II) , wherein X is halide, preferably chloride or bromide, most preferably chloride, with i) MOOH/M₂O₂, wherein M is any metal or ammonium-containing group that will react with H₂O₂ to form MOOH/M₂O₂ without decomposing one or more of the peroxides present in the process, 25 preferably M is selected from the group consisting of ammonium, sodium,

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potassium, magnesium, calcium, and lithium, and/or ii) one or more



peracids of formula (III) R^1 , or the M salts thereof, in an aqueous phase, in which process the acid halide, or mixture of acid halides, is essentially only brought into contact with water containing $MOOH/M_2O_2$ and/or one or more peracids or peracid salts, preferably M_2O_2 or the M salt of the peracid,

5 and/or one or more peracids or peracid salts, preferably M_2O_2 or the M salt of the peracid,

- optionally introducing one or more solvents for the acid halide in any part of the process,

- optionally introducing one or more salts in any part of the process,

10 - optionally introducing one or more colloids and/or surfactants, before, during, or after the peroxidation step,

- optionally purifying the aqueous mixture in one or more purification steps,

- optionally homogenizing the aqueous mixture in one or more homogenization steps,

15 - transferring the product from the previous steps that comprises the diacyl peroxide of formula (I) to a polymerization reactor, and

- thermally decomposing said diacyl peroxide, to generate organic free radicals, in the presence of one or more ethylenically unsaturated monomers, to polymerize said monomers in said polymerization reactor,

20 wherein the diacyl peroxide of formula (I), when transferred to the polymerization reactor, is always transferred from the previous step in the aqueous phase.

6. A process according to any one of the preceding claims wherein the diacyl peroxide obtained in the peroxidation step is used in the polymerization process within a period of 168 hours after said peroxidation step.

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7. A process according to any one of the preceding claims wherein the diacyl peroxide is stored and handled in this process at temperatures up to 5°C.
8. A process according to any one of the preceding claims wherein one and the
5 same reactor is used to make the MOOH/M₂O₂ from a source of M and H₂O₂ and to carry out the peroxidation step.
9. A process according to any one of the preceding claims wherein 0.005 to 2
10 %w/w of the diacyl peroxide is used in the polymerization step, based on the weight of the ethylenically unsaturated monomer(s).
10. A process according to any one of the preceding claims wherein the ethylenically unsaturated monomer(s) comprise(s) vinyl chloride monomer.
- 15 11. A process according to any one of the preceding claims wherein at least one acid halide is used that is selected from the group consisting of lauroyl halide, isobutanoyl halide, 2-ethylbutanoyl halide, 2-ethylhexanoyl halide, 2-chloropropanoyl halide, cyclohexane carbonyl halide, 3-methylcyclohexane carbonyl halide, and 2,3-dihalocyclohexane carbonyl halide.
- 20 12. A process according to any one of the preceding claims wherein at least one resulting diacyl peroxide is asymmetrical.
13. A process according to claim 12 wherein a peracid is used in the peroxidation
25 step.
14. A process according to any one of the preceding claims that is essentially solvent-free.

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15. A process according to any one of the preceding claims wherein the diacyl peroxide is dosed to the polymerization reaction mixture at the reaction temperature.

INTERNATIONAL SEARCH REPORT

International Application No
PCT/EP 03/01746

A. CLASSIFICATION OF SUBJECT MATTER
IPC 7 C08F4/34 C07C409/34

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
IPC 7 C08F C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category [°]	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	DE 10 68 014 B (I.C.I.) 29 October 1959 (1959-10-29) claim 1 ----	1,5-11
X	US 3 936 506 A (R. BERTHOLD) 3 February 1976 (1976-02-03) claim 1 ----	1,5-11
X	US 3 502 701 A (R. LEWIS) 24 March 1970 (1970-03-24) claim 1 -----	1,5-11

Further documents are listed in the continuation of box C.

Patent family members are listed in annex.

[°] Special categories of cited documents :

- *A* document defining the general state of the art which is not considered to be of particular relevance
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T later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

X document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

Y document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.

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Date of the actual completion of the international search

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INTERNATIONAL SEARCH REPORT

Information on patent family members

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