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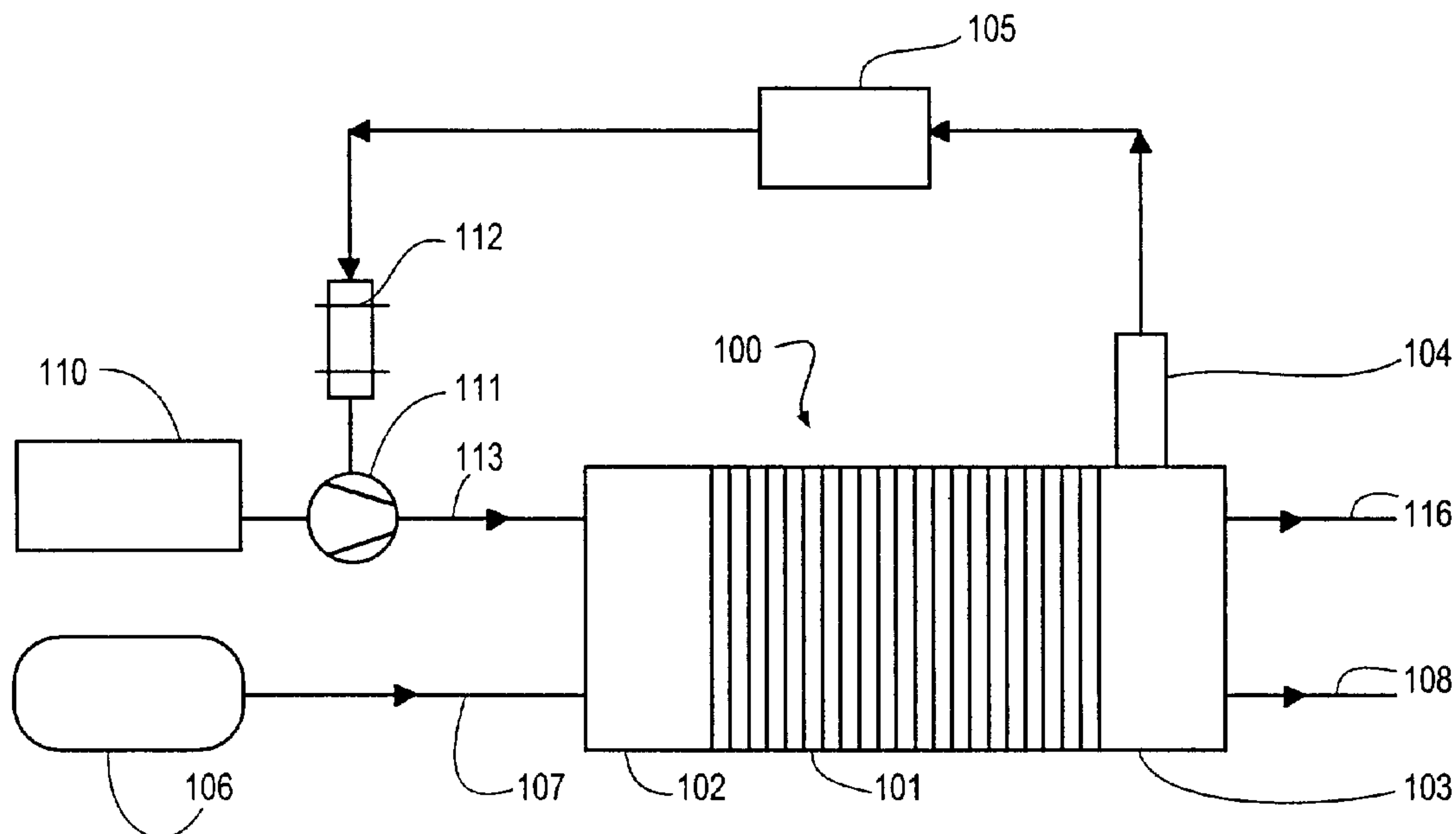
(71) Demandeur/Applicant:
BALLARD POWER SYSTEMS INC., CA

(72) Inventeurs/Inventors:
BOEHM, GUSTAV, DE;
WILKINSON, DAVID P., CA;
KNIGHTS, SHANNA, CA;
SCHAMM, REINHOLD, DE;
FLETCHER, NICHOLAS J., CA

(74) Agent: RUSSELL REYNEKE

(54) Titre : PROCEDE ET DISPOSITIF DE FONCTIONNEMENT D'UNE PILE A COMBUSTIBLE

(54) Title: METHOD AND APPARATUS FOR OPERATING A FUEL CELL



(57) Abrégé/Abstract:

The present invention relates to improving the overall efficiency of a fuel cell system by reducing parasitic power consumption. In particular, efficiency is improved by controlling the supply of oxidant to reduce excess oxidant stream flow by operating the fuel cell system with an oxidant stoichiometry between about one and two. A controller is programmed to decrease oxidant stoichiometry until oxidant starvation is detected or until oxidant stoichiometry is about one. When oxidant starvation is detected, the oxidant stoichiometry is increased until oxidant starvation is no longer detected. The fuel cell system employs a sensor for detecting an operational characteristic such as voltage output, or oxygen or hydrogen concentration in the cathode exhaust stream. The controller uses the operational characteristic to calculate oxidant stoichiometry or to determine when there is oxidant starvation at the cathode.

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(71) Applicant (for all designated States except US): **BALLARD POWER SYSTEMS INC.** [CA/CA]; 9000 Glenlyon Parkway, Burnaby, British Columbia V5J 5J9 (CA).

(72) Inventors; and

(75) Inventors/Applicants (for US only): **BOEHM, Gustav** [DE/DE]; Zur Forelle 34, 88662 Ueberlingen (DE). **WILKINSON, David, P.** [CA/CA]; 1391 Coleman Street, North Vancouver, British Columbia V7K 1W4 (CA). **KNIGHTS, Shanna** [CA/CA]; 5376 Forest Street,

Burnaby, British Columbia V5G 1X2 (CA). **SCHAMM, Reinhold** [DE/DE]; Muensterlinger weg 12, 88179 Stetten (DE). **FLETCHER, Nicholas, J.** [CA/CA]; 3464 Point Grey Road, Vancouver, British Columbia V6R 1A5 (CA).

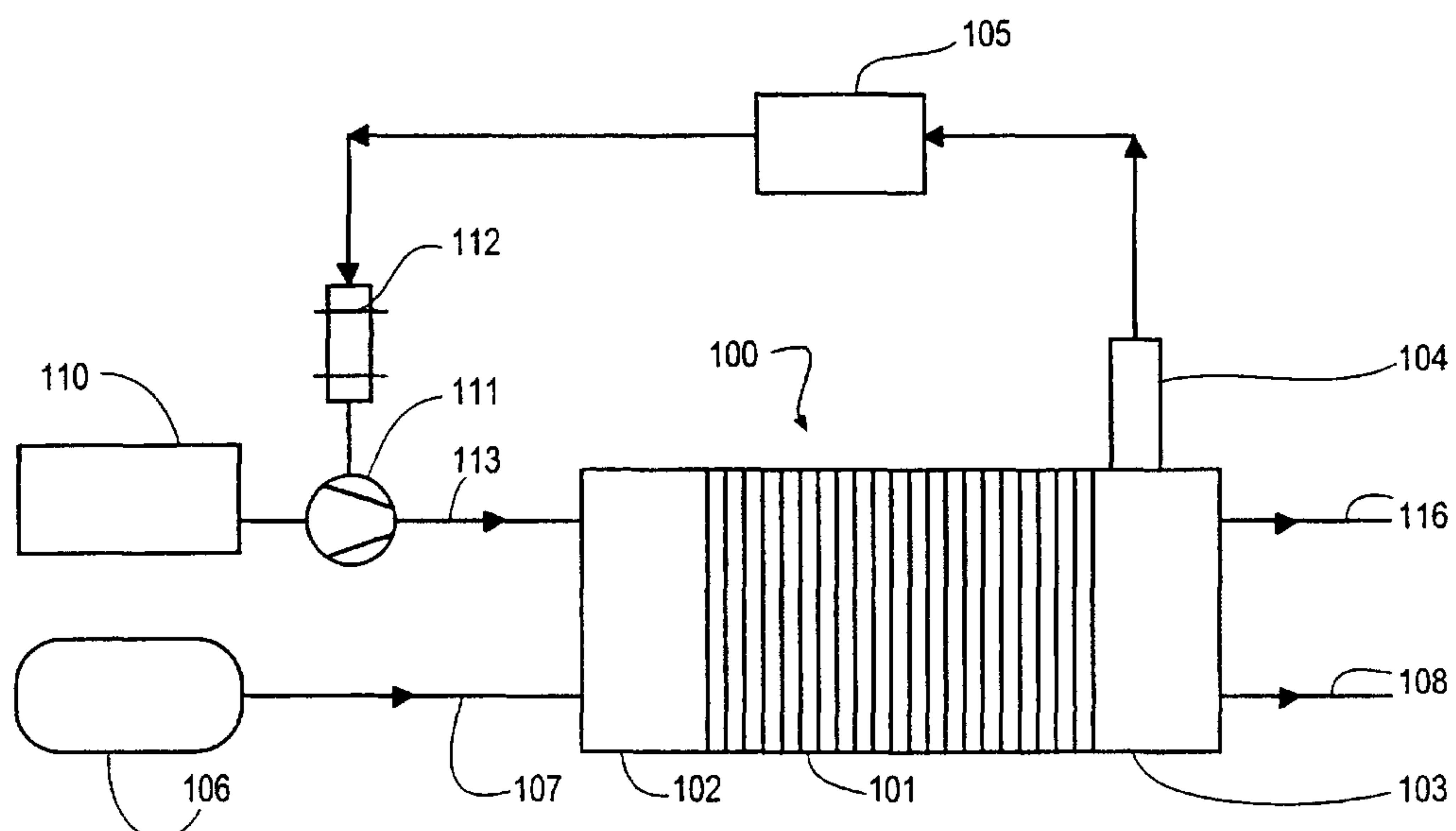
(74) Agent: **DE KOCK, Elbie, R.**; Russell Reyneke, Two Bentall Centre, Suite 700, 555 Burrard Street, Vancouver, British Columbia V7X 1M8 (CA).

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(54) Title: METHOD AND APPARATUS FOR OPERATING A FUEL CELL



(57) Abstract: The present invention relates to improving the overall efficiency of a fuel cell system by reducing parasitic power consumption. In particular, efficiency is improved by controlling the supply of oxidant to reduce excess oxidant stream flow by operating the fuel cell system with an oxidant stoichiometry between about one and two. A controller is programmed to decrease oxidant stoichiometry until oxidant starvation is detected or until oxidant stoichiometry is about one. When oxidant starvation is detected, the oxidant stoichiometry is increased until oxidant starvation is no longer detected. The fuel cell system employs a sensor for detecting an operational characteristic such as voltage output, or oxygen or hydrogen concentration in the cathode exhaust stream. The controller uses the operational characteristic to calculate oxidant stoichiometry or to determine when there is oxidant starvation at the cathode.



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METHOD AND APPARATUS FOR OPERATING A FUEL CELL

Field Of The Invention

10 The present invention relates to a method and apparatus for operating a fuel cell that improves the overall efficiency of the fuel cell system. In particular, efficiency is improved by controlling the supply of oxidant so as to reduce
15 excess oxidant flow.

Background Of The Invention

 Electrochemical fuel cells convert reactants, namely, fuel and oxidant fluid streams, to
20 generate electric power and reaction products. Electrochemical fuel cells generally employ an electrolyte disposed between two electrodes, namely a cathode and an anode. The electrodes each comprise an electrocatalyst disposed at the
25 interface between the electrolyte and the electrodes to induce the desired electrochemical reactions.

 The fuel fluid stream, which is supplied to the anode, typically comprises hydrogen and may be
30 pure gaseous hydrogen or a dilute hydrogen stream such as a reformat stream. Alternatively, other fuels such as methanol or dimethyl ether may be supplied to the anode where such fuels may be directly oxidized. The oxidant fluid stream,
35 which is supplied to the cathode, typically comprises oxygen, and may be pure gaseous oxygen, or a dilute oxygen stream such as air.

 For a fuel cell, reactant stoichiometry is defined herein as the ratio of the reactant
40 supplied over the reactant theoretically required

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5 to produce the current produced by the fuel cell.
For conventionally operated fuel cells which
typically supply a surplus of oxidant to the
cathode, since the oxidant is preferentially
reduced at the cathode, oxidant stoichiometry is
10 commonly expressed as the ratio of the oxidant
supplied over the oxidant consumed. However, at
lower stoichiometries, the reduction of oxidant
may not be responsible for all of the current
produced by the fuel cell. Other reactions, such
15 as, for example, the reduction of protons may also
occur at the cathode and contribute to the current
output (i.e. with the consequence of reduced
output voltage). In this example, while oxidant
may still be the main component being reduced at
20 the cathode, the amount of oxidant theoretically
required to produce the current output may be
higher than the amount of oxidant actually
supplied. Therefore, when components other than
the oxidant are reduced at the cathode, oxidant
25 stoichiometries less than one may be sustained.
If the oxidant is a dilute oxidant stream such as
air, only the reactant component, namely oxygen,
is considered in the calculation of stoichiometry
(that is, oxidant stoichiometry is the ratio of
30 the amount of oxygen supplied over the theoretical
amount of oxygen required to produce the fuel cell
output current).

Hydrogen and oxygen are reactive in the fuel
cell and are particularly reactive with each
35 other. Accordingly, in solid polymer fuel cells,
an important function of the membrane electrolyte
is to keep the hydrogen supplied to the anode

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5 separated from the oxygen supplied to the cathode.
In addition, the membrane is proton conductive and
functions as an electrolyte.

The overall efficiency of a fuel cell system
is a function of the total power output of the
10 fuel cell(s) and the parasitic power consumption.
Total parasitic power consumption is defined
herein as the sum of all power that is consumed by
the fuel cell system in the course of generating
electrical power. The net electrical power output
15 is the total power output minus the total
parasitic power consumption. Therefore, overall
efficiency may be improved by reducing the
parasitic power consumption.

One source of parasitic power consumption,
20 for example, is the oxidant delivery subsystem
that typically employs a mechanical device such as
a compressor, fan, pump, rotary piston blower, or
an equivalent mechanical device that consumes
power to supply oxidant to the fuel cell. Higher
25 oxidant stoichiometries generally result in higher
parasitic power consumption because more power is
generally required to deliver more oxidant to the
cathode. Conventional fuel cell systems typically
operate with an oxidant stoichiometry greater than
30 two (2.0). Since conventional fuel cell systems
direct at least twice the amount of oxygen to the
cathode than is actually required to satisfy the
electrical power demand, a significant amount of
the parasitic power consumption is for directing
35 surplus oxygen to the cathode. Further, fuel cell
systems commonly employ a dilute oxidant stream.
A dilute oxidant stream is defined herein as a

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5 fluid stream that comprises less than 100%
oxidant. For example, air is a dilute oxidant
stream that typically comprises about 20% oxygen,
in addition to other components such as nitrogen.
Accordingly, when air is employed as the dilute
10 oxidant stream, parasitic power consumption is
amplified because in addition to the surplus
oxygen, the oxidant delivery subsystem must also
supply a proportionate amount of the other non-
reactive components.

15 One reason why the parasitic power
consumption associated with high oxidant
stoichiometries is tolerated, is that excess
oxidant is desired at the cathode to avoid oxidant
starvation at the cathode electrocatalyst.
20 Oxidant starvation is defined herein as the status
when oxidant stoichiometry is less than one.
Oxygen starvation typically results in a
condition, in the absence of oxidant at the
cathode electrocatalyst, favoring the production
25 of molecular hydrogen from protons and electrons
at the cathode. In severe cases of oxidant
starvation the fuel cell may generate a negative
voltage and this condition is known as cell
reversal. Oxidant stream delivery systems are
30 typically designed to provide a generous surplus
of oxidant to maintain performance, to reduce the
likelihood of oxidant starvation, and to reduce
the likelihood of hydrogen production at the
cathode, even though this results in the
35 aforementioned amplified parasitic power
consumption.

In fuel cells, oxidant starvation is most

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5 likely to occur in regions furthest downstream
from the cathode inlet where the oxidant stream
enters the cell, for example, near the cathode
outlet. An oxidant stoichiometry that provides a
surplus of oxidant to the fuel cell provides an
10 adequate concentration of oxidant to the
electrocatalyst throughout the electrochemically
active area of the cathode, including near the
cathode outlet.

Another reason why conventional fuel cell
15 systems seek to avoid low oxidant stoichiometries
is that the temperature within the fuel cell may
rapidly increase when oxidant stoichiometry is too
low. It is generally desirable to maintain the
temperature of solid polymer fuel cells below
20 100°C. When the temperature increases within the
fuel cell, parasitic power consumption increases
because of the higher load on the cooling system,
offsetting, to some degree, the reduction in
parasitic power consumption associated with
25 operating at a lower oxidant stoichiometry.
Another disadvantage of operating a fuel cell
system with a high oxidant stoichiometry is that
higher oxidant stoichiometries generally require
higher speeds for the mechanical devices used by
30 the oxidant delivery subsystem to supply the
oxidant stream to the cathode. Now that fuel cell
systems are being developed for commercial use,
mechanical considerations over the planned
lifetime of commercial fuel cell systems are a
35 factor. A mechanical disadvantage of conventional
high stoichiometry methods of operation is that
such methods may result in increased wear and more

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frequent maintenance. If the oxidant is air, there may be additional operational costs because supplying a high surplus of oxidant also results in higher flow rates that may increase air filter
5 maintenance and/or reduce filter efficiency.

Summary of the Invention

A method of operating a fuel cell system controls oxidant stoichiometry to reduce parasitic
10 power consumption to improve overall efficiency, while avoiding low oxidant stoichiometries that might cause reduced performance, cell reversal, hydrogen production at the cathode, and increased heat generation within the fuel cell. The fuel
15 cell system comprises a fuel cell power generating subsystem having at least one fuel cell, and an oxidant delivery subsystem that comprises at least one mechanical device for supplying an oxidant stream to a cathode of the fuel cell. The fuel
20 cell also has an anode supplied with a fuel stream. In a preferred embodiment, the fuel cell is a solid polymer fuel cell.

The method comprises controlling the mechanical device, to reduce parasitic power
25 consumption by reducing the oxidant stoichiometry until $dV/d(OS)$ is greater than a predetermined value ("PV"), where dV is the change in cell voltage and $d(OS)$ is the change in oxidant stoichiometry (that is, the slope of a plot of
30 voltage as a function of oxidant stoichiometry). Cell voltage is measured in volts and oxidant stoichiometry is a unit-less ratio.

To practice the invention, the value of $dV/d(OS)$ need not actually be calculated if a

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5 $dV/d(OS)$ need not actually be calculated if a
relationship between $dV/d(OS)$ and another
operational characteristic is known. For example,
in preferred embodiments, an operational
characteristic that correlates to $dV/d(OS)$ and/or
10 oxidant stoichiometry may be monitored. The fuel
cell system is controlled to take action when the
value of the monitored operational characteristic
correlates to when $dV/d(OS)$ is equal to or greater
than PV. For example, in a typical fuel cell
15 system, during normal operation, current density
is kept constant and when oxidant stoichiometry is
being reduced, a particular cell voltage
correlates to when $dV/d(OS)$ increases to PV. That
is, when cell voltage decreases below a threshold
20 voltage, this is determined when $dV/d(OS)$ is
higher than PV. Accordingly, a fuel cell system
may be operated to reduce parasitic power
consumption by controlling the oxidant delivery
subsystem to maintain voltage output within a
25 predetermined voltage range which typically
corresponds to an oxidant stoichiometry range
between about one and two, wherein PV is selected
so that $dV/d(OS)$ equals PV at the lower limit of
the selected voltage range. The preferred range
30 for oxidant stoichiometry may actually change
according to the instantaneous operating
conditions. For example, when a fuel cell is
operating in an idle or low output condition, a
higher oxidant stoichiometry may be preferred to
35 prevent accumulation of water at the cathode.
Accordingly, the value of PV may be dynamic.

The characteristics of the fuel cell and/or

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5 the type of reactants may also influence the preferred oxidant stoichiometry range. For example, in a direct methanol fuel cell, higher stoichiometries are typically employed, but the reactant supply may still be controlled to prevent
10 $dV/d(OS)$ from increasing to higher than PV (although for a direct methanol fuel cell, PV will correspond to a higher stoichiometry, compared to a fuel cell which is fed hydrogen gas or reformat as the fuel stream).

15 Similarly, when current density is constant and oxidant stoichiometry is being reduced, a particular oxidant stoichiometry correlates to when $dV/d(OS)$ increases to PV. Accordingly, operational characteristics, such as the oxygen
20 concentration in the cathode exhaust stream, which correlate to oxidant stoichiometry, may be monitored to determine when oxidant stoichiometry is reduced to a value which correlates to when $dV/d(OS)$ increases to greater than or equal to PV.
25 The oxidant concentration in the oxidant supply stream is typically known, but if the oxidant supply stream has a variable oxidant concentration (for example, if an oxidant enrichment system is employed), the method may further comprise
30 monitoring and measuring the oxidant concentration in the oxidant supply stream, in addition to monitoring and measuring the oxidant concentration in the oxidant exhaust stream. Alternatively, oxidant stoichiometry may be determined by
35 monitoring and measuring a different operational characteristic, such as, for example, current output for the fuel cell power generating

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5 subsystem, which, in addition to the oxidant concentration in the oxidant exhaust stream may be used to calculate oxidant stoichiometry.

The value of $dV/d(OS)$ generally increases as oxidant stoichiometry and cell voltage both
10 decrease. In one embodiment, PV corresponds to when oxidant starvation is beginning to occur or when oxidant starvation is beginning to cause a decline in performance. In a more preferable embodiment, PV corresponds to when further
15 reductions in oxidant stoichiometry will cause a sharp decline in cell voltage output, for example, when $dV/d(OS)$ is higher than 0.02 volts.

Preferably, PV is between 0.3 volt and 7.0 volts, so that the fuel cell system operates mostly when
20 $dV/d(OS)$ is less than PV. The selected value for PV controls the oxidant stoichiometry so that it is kept between about one and two during normal operation and closer to about one or a predetermined target value, preferably between 1
25 and 1.5, during steady state operation.

In a preferred apparatus for practising the method, the fuel cell is one of a plurality of fuel cells arranged in a stack. When the method is applied to a fuel cell stack, the sensor may
30 monitor the operational characteristic for one or more individual fuel cells and/or for the stack as a whole. The sensor may thus be located to monitor an operational characteristic (for example, oxidant or hydrogen concentration) within
35 a portion of a reactant passage (for example, an internal cathode exhaust passage) that is disposed between the outside end surfaces of the stack end

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5 plates.

The oxidant stoichiometry is preferably controlled by controlling the oxidant stream mass flow rate, for example, by controlling the speed of a mechanical device, such as a compressor, a fan, a pump, or a blower. Reducing the speed of the mechanical device generally reduces parasitic power consumption and reduces oxidant stoichiometry. However, alternate methods of controlling oxidant stoichiometry may be employed which also reduce parasitic power consumption. For example, if an oxidant enrichment subsystem is employed, oxidant stoichiometry may be controlled by increasing or decreasing the concentration of oxidant in the oxidant stream supplied to the cathode(s) of the fuel cell power generating subsystem. Another method of controlling oxidant stoichiometry is adjusting the electrical power output of the fuel cell, wherein reducing power output generally increases oxidant stoichiometry.

25 A preferred method that employs a hydrogen sensor (the "Hydrogen Sensor Method") comprises:

- (a) monitoring a cathode exhaust stream downstream of the cathode to detect hydrogen gas concentration; and
- 30 (b) decreasing oxidant stoichiometry when the hydrogen gas concentration is less than a first threshold concentration.

The Hydrogen Sensor Method may further comprise increasing the oxidant stoichiometry when the hydrogen concentration is higher than a second threshold concentration (for example, 20 ppm of hydrogen), which correlates to operating

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5 conditions which are indicative of actual or
potential oxidant starvation. The first threshold
concentration may be, for example, the lower
detection limit of the hydrogen sensor that is
used to monitor the cathode exhaust stream. The
10 second threshold concentration is greater than the
first threshold concentration. When the hydrogen
concentration is between the first and second
threshold concentrations, the controller does not
take any action to adjust the oxidant
15 stoichiometry.

A problem with using the hydrogen
concentration measured in the cathode exhaust
stream to detect oxidant starvation is that
oxidant starvation is not the only possible cause
20 for hydrogen gas being detected at the cathode.
For example, when the fuel comprises hydrogen,
holes or cracks may form in the membrane or seals
and permit reactants to "cross over" from the
anode side to the cathode side, and vice versa.
25 If significant reactant crossover is detected, the
conventional response is to shut down the fuel
cell so that it may be repaired or replaced. Fuel
crossover and oxidant starvation may both cause
reduced fuel cell performance, but the detection
30 of one condition requires a response which is
different from the response required for the other
condition. Oxidant starvation causing hydrogen to
be produced at the cathode generally requires the
oxidant stoichiometry to be increased, whereas
35 fuel crossover, if significant, may require the
fuel cell to be shut down. Therefore, for
appropriate action to be taken, it is desirable

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5 for the controller to be able to distinguish
between oxidant starvation and fuel crossover when
the fuel comprises hydrogen. The following
embodiments of the Hydrogen Sensor Method provide
procedures for distinguishing between oxidant
10 starvation and fuel crossover.

 The Hydrogen Sensor Method may further
comprise steps for reducing the hydrogen gas
concentration within the cathode exhaust stream
when the hydrogen gas concentration is greater
15 than a second threshold concentration, wherein the
steps comprise comparing the oxidant stream mass
flow rate to a maximum desired mass flow rate, and

 (a) if the oxidant stream mass flow rate is
less than the maximum desired mass flow
20 rate, increasing the oxidant mass flow
rate (that is, if raising the oxidant
mass flow rate results in less hydrogen
being detected at the cathode, then it
is confirmed that oxidant starvation was
25 likely the reason for hydrogen being
detected; if oxidant starvation is not
the source of the hydrogen at the
cathode, the oxidant stream mass flow
rate will quickly increase to the
30 maximum desired mass flow rate and the
controller will determine that fuel
crossover is the likely hydrogen
source); and

 (b) if the oxidant mass flow rate is already
35 greater than or equal to the maximum
desired mass flow rate,
ceasing operation of the fuel cell if

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the hydrogen gas concentration is greater than a third concentration threshold which is greater than the first and second concentration thresholds (that is, since the
5 oxidant stream mass flow rate is already at, or exceeds, the desired maximum, oxidant starvation is not the likely source of the hydrogen in the cathode exhaust stream; since the hydrogen concentration is above the third
10 threshold, this indicates that there may be an excessive amount fuel passing through leaks between the anode and cathode); and generating a warning signal and continuing to operate the fuel cell if the
15 hydrogen gas concentration is less than the third concentration threshold (that is, the value of the third threshold is selected so that the fuel cell system can be safely operated when the hydrogen concentration in
20 the cathode exhaust stream is less than the third threshold).

The Hydrogen Sensor Method may further comprise continuously monitoring the cathode exhaust stream for the hydrogen gas concentration
25 and determining whether the hydrogen gas concentration is increasing or decreasing, and when the hydrogen gas concentration is greater than a second threshold concentration, the method further comprises:

30 maintaining a substantially constant oxidant stoichiometry when the hydrogen concentration is decreasing; and increasing the oxidant stoichiometry

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when the hydrogen concentration is increasing.

In addition to monitoring whether hydrogen concentration is increasing or decreasing, when
5 the hydrogen gas concentration is greater than a second threshold concentration, the method may also comprise additional steps to determine whether the source of the hydrogen is oxidant starvation or fuel crossover. For example, the
10 additional steps may comprise:

measuring fuel cell voltage and comparing it to a voltage threshold value (for a Ballard® MK V fuel cell, the voltage threshold value could be, for example, 100
15 millivolts), and

if the fuel cell voltage exceeds the voltage threshold value and the hydrogen gas concentration is increasing, decreasing the pressure of
20 the fuel stream (in this case, since voltage exceeds the threshold value, the reason for the increasing hydrogen concentration is probably a leak; to reduce the effect of the leak, the
25 method preferably comprises controlling the fuel stream pressure so that it is less than or equal to the pressure of the oxidant stream);

if the fuel cell voltage is less
30 than the voltage threshold value, the hydrogen gas concentration is increasing, and oxidant mass flow rate is less than a desired maximum, then

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5 increasing the oxidant stoichiometry (in
this case, since oxidant mass flow rate
is less than the desired maximum, the
cause for the low cell voltage and the
presence of hydrogen gas may be oxidant
10 starvation, and the controller attempts
to correct this condition by increasing
the oxidant stoichiometry); and

 if the fuel cell voltage is less
than the voltage threshold value, the
15 hydrogen gas concentration is
increasing, and oxidant mass flow rate
is greater than or equal to a desired
maximum, then decreasing the pressure of
the fuel stream (in this case, the low
20 cell voltage may be caused by oxidant
starvation or fuel leaking from the
anode to the cathode; since the oxidant
mass flow rate is already greater than
or equal to the desired maximum, the
25 pressure of the fuel stream is reduced,
thereby reducing fuel cell power output
and oxygen consumption at the cathode,
to counter oxidant starvation and reduce
the effect of any leaks).

30 Further additional steps may be taken to
confirm whether the detected hydrogen gas
concentration is caused by oxidant starvation or
fuel crossover. For example, the method may also
comprise regulating fluid pressure of the oxidant
and fuel streams to increase or decrease a
35 pressure differential between the oxidant and fuel
streams to help determine whether the hydrogen

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5 measured at the cathode is caused by a leak or by
oxidant starvation. If the change in the pressure
differential has a significant effect on the
measured hydrogen concentration, then it can be
determined that there is a significant problem
10 with hydrogen crossover.

For any of the above-described methods, the
oxidant stoichiometry is typically adjusted by
controlling the speed of the oxidant compressor or
blower. However, other methods of changing the
15 oxidant stoichiometry may also be used, such as
adjusting the oxidant concentration in the oxidant
supply stream or changing the electrical power
output of the fuel cell without changing the mass
flow rate of the oxidant supply stream. When the
20 oxidant stream mass flow rate is adjusted, it is
typically changed by a fixed amount or by a fixed
percentage of the instant oxidant stream mass flow
rate. Alternatively, oxidant stoichiometry may be
adjusted by adjusting the oxidant stream mass flow
25 rate by an amount that is dependent upon the
magnitude of the detected hydrogen gas
concentration. For example, the controller may be
programmed to reduce the oxidant stoichiometry by
a larger amount when a large surplus of oxygen is
30 detected compared to when only a small surplus of
oxygen is detected.

The method of controlling the oxidant
delivery subsystem to reduce parasitic power
consumption may comprise calibrating an oxidant
35 delivery subsystem for a fuel cell. For example,
the calibration method may comprise:

(a) operating the fuel cell at a particular

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- 5 electrical power output;
- (b) supplying an oxidant stream to a cathode
of the fuel cell;
- (c) adjusting the operating speed of a
mechanical oxidant delivery device;
- 10 (d) measuring an operational characteristic
that corresponds to $dV/d(\text{oxidant stoichiometry})$; and
- (e) recording as the desired operating speed
for said particular electrical power
15 output, said operating speed when said
 $dV/d(\text{oxidant stoichiometry})$ is equal to
a predetermined value.

The calibration method may be repeated for a
plurality of electrical power outputs so that the
20 desired operating speed for the mechanical oxidant
delivery device may be determined and recorded in
a look-up table for many different electrical
power demands. The desired operating speed may
then be determined by referring to a look-up table
25 for the operating speed that corresponds to the
instant electrical power demand.

An advantage of the calibration method is
that it may be used throughout the operating life
of the fuel cell to adjust for changes in the fuel
30 cell over time. For example some of the fuel cell
properties may be subject to degradation over time
and that may change the stoichiometry requirements
over the operational lifetime of the fuel cell.

The present method and apparatus also
35 controls the amount of oxidant supplied to a fuel
cell stack and reduces system inefficiencies
caused by the over-supply of oxidant. Preferably,

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5 the method also controls the oxidant delivery
subsystem to increase the oxidant stoichiometry to
avert oxidant starvation conditions. Accordingly,
a method of operating a fuel cell is provided that
detects when the oxidant stoichiometry may be
10 decreased or increased, and when the flow of
oxidant should be discontinued altogether.

The present method may also be employed to
operate a fuel cell and further reduce parasitic
power consumption by controlling the supply of
15 fuel to reduce excess fuel flow. The same
principles that apply to the present method for
controlling the oxidant supply apply to a method
for controlling the fuel supply. Fuel cells
typically employ a mechanical device, such as, for
20 example, a compressor or pump, to supply a fuel
stream to the anode(s) of the fuel cell(s).
Therefore, parasitic power consumption may be
reduced by reducing fuel stoichiometry to reduce
the amount of excess fuel supplied to the fuel
25 cell anode(s) and the work performed by the
compressor. A reduction of the fuel stoichiometry
generally causes an increase in $dV/d(\text{fuel stoichiometry})$. According to the present method,
fuel stoichiometry is kept within a predetermined
30 range by reducing fuel stoichiometry until
 $dV/d(\text{fuel stoichiometry})$ increases above a
predetermined threshold value. The predetermined
range and threshold value depend upon the
particular characteristics and operating
35 conditions of each particular fuel cell or fuel
cell stack. The predetermined range may be
empirically determined, for example, with

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5 consideration to these factors.

Generally, it is desirable to reduce reactant stoichiometry until $dV/d(\text{reactant stoichiometry})$ is greater than about 0.02 volt. More preferably, the predetermined value for $dV/d(\text{reactant}$
10 $\text{stoichiometry})$ is between 0.30 and 7.0 volts. The voltage drop is generally more severe with fuel starvation, compared to oxidant starvation and when cell voltage is one of the monitored characteristics, this effect may be used to help
15 differentiate between oxidant or fuel starvation.

Brief Description Of The Drawings

The advantages, nature and additional features of the invention will become more
20 apparent from the following description, together with the accompanying drawings, in which:

FIG. 1 is a schematic diagram of a fuel cell system comprising a detector for detecting oxidant starvation at the cathode and a controller for
25 processing information from the detector and controlling the oxidant delivery subsystem to increase or decrease the oxidant stoichiometry;

FIGs. 2a, 2b and 2c are plots of experimental data that illustrate the effect oxidant
30 stoichiometry has on operational characteristics such as voltage output, heat produced, and $dV/d(\text{oxidant stoichiometry})$. FIG. 2a is a plot of fuel cell voltage output as a function of oxidant stoichiometry for a solid polymer fuel cell
35 operating at a current density of 500 amps per square foot (about 540 milliamps per square centimeter), during normal operation. FIG. 2b is a

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5 plot of fuel cell voltage output and $dV/d(\text{oxidant stoichiometry})$ as a function of oxidant stoichiometry for the same fuel cell experiment as FIG. 2a. FIG. 2c is a plot of $dV/d(\text{oxidant stoichiometry})$ as a function of oxidant
10 stoichiometry for a fuel cell stack comprising four fuel cells;

FIGs. 3a, 3b and 4-12 are logic diagrams illustrating various preferred embodiments of a the present method of operating a fuel cell and
15 controlling the oxidant stoichiometry. In FIGs. 3a, 3b and 4-11, the logic diagrams illustrate a method of adjusting the oxidant stream mass flow rate to adjust the oxidant stoichiometry. In these embodiments, the oxidant stream mass flow
20 rate may also be changed in response to changes in power output, with the illustrated method being used to adjust the oxidant stream mass flow rate and hence oxidant stoichiometry to prevent oxidant starvation and reduce parasitic power losses
25 associated with supplying excess oxidant. In FIG. 12 the logic diagram illustrates a method of calibrating a fuel cell system to determine the desired oxidant stream mass flow rate to produce a particular power output at a predetermined oxidant
30 stoichiometry.

Detailed Description Of Preferred Embodiments

FIG. 1 is a schematic diagram that shows the fuel cell power generating subsystem, the oxidant
35 delivery subsystem, and the fuel delivery subsystem of a fuel cell system. The fuel cell power generating subsystem comprises a fuel cell

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5 stack 100, comprising a plurality of fuel cells
101 interposed between end plates 102 and 103.
The fuel cell power generating subsystem further
comprises sensor 104. When fuel cell stack 100 is
operating, sensor 104 measures an operational
10 characteristic that correlates to $dV/d(OS)$. For
example, when fuel cell stack 100 is operating at
a constant current density, sensor 104 may measure
an operational characteristic which relates to
oxidant stoichiometry, cell voltage, or a
15 characteristic which is typically detected when
oxidant starvation is occurring at the fuel cell
cathode.

Sensor 104 outputs a signal to controller 105
which processes the signal to determine when
20 $dV/d(OS)$ is within the desired operating range and
when oxidant stoichiometry should be adjusted so
that $dV/d(OS)$ is restored to the desired operating
range. For example, oxidant stoichiometry may be
increased if the measured operational
25 characteristic indicates an oxidant stoichiometry
and/or the presence of conditions at the cathode
which indicates actual, or a potential for,
oxidant starvation. Preferably, the desired
operating range prevents any oxidant starvation at
30 the cathode that inhibits the fuel cell from
producing the desired power output.

When the operational characteristic monitored
by sensor 104 is the concentration of a gas in the
cathode exhaust stream, sensor 104 may comprise a
35 sensing element that is located within the
interior of the cathode exhaust passage so that it
is exposed to the cathode exhaust stream. The

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5 portion of the cathode exhaust passage where the sensing element is located may be a manifold or fluid passage internal to fuel cell stack 100, or in cathode exhaust passage 116.

10 In one embodiment, oxidant stoichiometry is controlled by using sensor 104 to measure an operational characteristic that correlates to oxidant stoichiometry. In another embodiment, sensor 104 detects an operational characteristic that is indicative of oxidant starvation at the
15 cathode so that the oxidant stoichiometry may be controlled to reduce the amount of excess oxidant supplied to fuel cell stack 100, while preventing harmful oxidant starvation at the fuel cell cathodes (that is, in this embodiment controller
20 105 increases oxidant stoichiometry when sensor 104 detects actual or potential oxidant starvation and may decrease oxidant stoichiometry when oxidant starvation is not detected). In yet another embodiment, controller 105 checks for
25 oxidant starvation while maintaining $dV/d(OS)$ within a predetermined operating range; if oxidant starvation is detected, oxidant stoichiometry is increased until oxidant starvation is no longer detected, even though this may result in
30 temporarily raising $dV/d(OS)$ above the desired operating range.

The fuel delivery subsystem supplies a fuel stream from fuel supply 106 to the anodes of fuel cell stack 100 via fuel supply passage 107. When
35 the fuel stream is a compressed gas, such as substantially pure hydrogen, fuel supply 106 may comprise a pressure vessel and a pressure control

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5 valve (not shown) for regulating the pressure of
the fuel stream supplied to fuel cell stack 100.
Alternatively, the fuel may be a liquid fuel such
as methanol and fuel supply 106 may comprise a
fuel tank. Liquid fuel may be supplied directly
10 to fuel cell stack 100 (that is, a so called
"liquid feed fuel cell"). Alternatively, fuels
such as methanol, natural gas, or other
hydrocarbons may be further processed to produce a
gaseous hydrogen-containing reformat stream, in
15 which case fuel supply 106 further comprises a
fuel processor. When the fuel storage tank is not
pressurized, the fuel delivery subsystem may
further comprise a compressor or pump for
controlling the pressure and mass flow rate of the
20 fuel stream supplied to fuel cell stack 100.
After the fuel stream has been directed to the
anodes of fuel cell stack 100 to participate in
the desired electrochemical reactions, a fuel-
depleted fuel exhaust stream is exhausted from
25 fuel cell stack 100 via fuel exhaust passage 108.

In the preferred embodiment illustrated by
FIG. 1, the oxidant delivery subsystem comprises
oxidant supply 110, mechanical device 111 for
raising the pressure of the oxidant supply stream,
30 and electric motor 112 coupled to mechanical
device 111 for providing power thereto. Oxidant
supply 110 may comprise a vessel for holding a
supply of oxidant, but more typically, oxidant
supply 110 comprises an air intake for receiving
35 and filtering air from the surrounding atmosphere.
From oxidant supply 110, the oxidant supply stream
is directed to mechanical device 111, which

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5 raises the pressure of the oxidant stream. The pressurized oxidant supply stream is directed to the fuel cell power generating subsystem via oxidant supply passage 113.

10 Controller 105 receives an output signal from sensor 104. The output signal is processed by controller 105 to determine whether $dV/d(OS)$ is within the desired operating range. Controller 105 communicates with the oxidant delivery subsystem to control the output of mechanical
15 device 111 to maintain $dV/d(OS)$ within a predetermined desired operating range (which preferably corresponds to an oxidant stoichiometry between about one and two).

For example, in the embodiment illustrated in
20 FIG. 1, controller 105 controls electric motor 112 to control the speed of mechanical device 111. Mechanical device 111 is typically a compressor such as a rotary piston compressor or a reciprocating piston compressor. However, other
25 types of mechanical devices may also be employed such as, for example a pump, a fan, or a blower. Mechanical device 111 raises the pressure of the oxidant supply stream to provide sufficient energy for directing the desired oxidant mass flow rate
30 to the fuel cell cathodes within fuel cell stack 100. After the cathodes, the oxygen-depleted oxidant stream is ultimately exhausted from fuel cell stack 100 through cathode exhaust passage 116.

35 In a preferred method, during steady state operation, controller 105 controls the oxidant delivery subsystem so that the value of $dV/d(OS)$

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corresponds to operating conditions when oxidant stoichiometry is close to about one. Steady state operation is defined herein as an operational mode for the fuel cell system when the power output of fuel cell stack 100 is substantially constant. During normal operation, when the power demand is dynamic, controller 105 may allow $dV/d(OS)$ to vary within a predetermined desired operating range that preferably corresponds to when oxidant stoichiometry is between about one and two. Normal operation is defined herein to exclude start-up and shut-down modes when controller 105 may allow a value for $dV/d(OS)$ that corresponds to higher or lower oxidant stoichiometries, respectively. According to a preferred method, parasitic power demands are reduced during normal operation by reducing the power consumption of the oxidant delivery subsystem, by keeping oxidant stoichiometry less than two and preferably close to about one during steady state operation.

In a first preferred embodiment, sensor 104 measures the voltage output from fuel cell stack 100. During normal operation, at constant current density, fuel cell voltage output correlates to oxidant stoichiometry, so if sensor 104 measures fuel cell voltage output, sensor 104 may be used to determine oxidant stoichiometry and $dV/d(OS)$. For example, plot A of FIG. 2a sets forth the voltage output (left y-axis) as a function of oxidant stoichiometry (x-axis) for a fuel cell operating at a constant current density of 500 amps per square foot (about 540 milliamps per square centimeter). That is, if at least one of cell voltage output or oxidant stoichiometry is known, then $dV/d(OS)$ may be determined by

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cell voltage output or oxidant stoichiometry is known, then $dV/d(OS)$ may be determined by referring to plot A. Plot C of FIG. 2b sets forth voltage output as a function of oxidant stoichiometry, and plot D of FIG. 2b sets forth $dV/d(OS)$ as a function of oxidant stoichiometry. Plot D of FIG. 2b shows that by selecting a predetermined threshold value for $dV/d(OS)$, oxidant stoichiometry may be reduced until it approaches one, without significant sacrifices to performance, as long as oxidant stoichiometry is increased when $dV/d(OS)$ is greater than or equal to the predetermined threshold value. For example, for the fuel cell of FIG. 2b, a predetermined threshold value for $dV/d(OS)$ could be a value between 0.02 and 0.3. For a threshold value of 0.3, oxidant stoichiometry could be reduced to about 1.2 before $dV/d(OS)$ would increase to above about 0.3; at this point voltage output is still higher than 0.6 volts so fuel cell performance is not significantly compromised.

In the example of FIGs. 2a and 2b, the fuel cell was a Ballard® MK V fuel cell which employed a solid polymer ion exchange membrane made from Nafion™ 117 (a copolymer of tetrafluoroethylene and perfluorovinylether sulfonic acid). The electrodes were made from carbon fiber paper with a thickness of 0.09 inch (about 2.29 mm) obtained from Toray Industries Inc. The catalyst layer on the electrodes was platinum black catalyst mixed with a tetrafluoroethylene binder. The catalyst loading on each electrode was 4 mg/cm².

In the example of FIGs. 2a and 2b, if fuel

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cell voltage output is monitored, $dV/d(OS)$ itself need not be monitored, since the relationship between voltage output and $dV/d(OS)$ may be determined (see, for example, FIG. 2b).

5 Accordingly, the oxidant delivery subsystem may be controlled to maintain voltage output between about 0.63 volt and about 0.67 volt, then oxidant stoichiometry will be maintained between about 1.2 and about 1.4. That is, if sensor 104 detects a
10 voltage output higher than 0.67 volt, controller 105 controls the oxidant delivery subsystem to reduce the speed of mechanical device 111, thereby reducing parasitic power consumption, reducing oxidant stoichiometry, and keeping $dV/d(OS)$ within
15 the desired predetermined range. If sensor 104 detects a voltage output less than 0.63 volts, then controller 105 controls the oxidant delivery subsystem to increase the speed of mechanical device 111, to lower $dV/d(OS)$ and increase voltage
20 output and oxidant stoichiometry, thereby preventing oxidant starvation at the fuel cell cathodes.

Persons skilled in the art will recognize that fuel cell stacks, with different features,
25 such as, for example, the size of the electrochemically active area, may operate under the same conditions and produce different voltages from those shown in FIG. 2a. However, for any particular fuel cell or fuel cell stack, a similar
30 relationship between voltage output and oxidant stoichiometry can be plotted and used during normal operating conditions to control oxidant stoichiometry to reduce parasitic power

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5 consumption and maintain $dV/d(OS)$ within a desired predetermined range.

FIG. 2a also shows plot B which sets forth, for the same fuel cell, theoretical heat produced within the fuel cell (right y-axis) as a function of oxidant stoichiometry (x-axis). The
10 theoretical amount of heat produced within the fuel cell was calculated by calculating the heat balance for the fuel cell. That is, the calculation determined the heat produced by
15 considering the total enthalpy of the inlet and outlet fluid streams and the power produced. Plot B shows that for this particular fuel cell there is a substantial increase in the heat produced therein when the oxidant stoichiometry is between 1.2 and 0.9. Because increases in temperature
20 within a fuel cell may cause increases in parasitic power consumption by the cooling subsystem, it is important to also consider this effect when selecting the desired predetermined
25 range for $dV/d(OS)$. Accordingly, for some fuel cells the desired range for $dV/d(OS)$ may correlate to an oxidant stoichiometry range between about 1.2 and 2.0.

FIG. 2c is a plot of the data from four fuel
30 cells that were arranged in a stack, with each fuel cell operating at a constant current density of 500 amps per square foot (about 540 milliamps per square centimeter). Plots E through H each set forth $dV/d(OS)$ as a function of oxidant
35 stoichiometry for a respective one of the four fuel cells in the stack. Plot I sets forth the average $dV/d(OS)$ against oxidant stoichiometry.

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5 FIG. 2c, like FIG. 2b, shows the relationship between oxidant stoichiometry and $dV/d(OS)$, and how $dV/d(OS)$ progressively increases as oxidant stoichiometry is reduced towards one.

FIG. 2c shows that, within a fuel cell stack,
10 there may be different values for $dV/d(OS)$ for different fuel cells. In this case, the fuel cell may be controlled with reference to the average $dV/d(OS)$. Alternatively, the value of $dV/d(OS)$ for a selected fuel cell may be monitored to
15 control the oxidant stoichiometry for an entire fuel cell stack. The selected fuel cell may be made more responsive to changes in oxidant stoichiometry so that oxidant stoichiometry may be controlled to prevent large fluctuations in fuel
20 cell performance. For example, the selected fuel cell may be designed so that its voltage output decreases more rapidly than the voltage output of the other fuel cells in the stack, so that oxidant stoichiometry may be increased before there is a
25 significant decrease in the voltage output of the fuel cell stack.

The logic diagrams of FIGs. 3a, 3b, and 4-12 will be explained with reference to the components of the fuel cell system shown in FIG. 1.

30 In a preferred embodiment, controller 105 may be programmed to perform the method shown in the logic diagram of FIG. 3a. In this method, controller 105 controls oxidant stoichiometry to reduce parasitic power consumption in response to
35 signals emitted from sensor 104, by controlling oxidant stream mass flow rate using motor 112 and mechanical device 111. The method starts at step

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5 120. In step 122 reactants are supplied to fuel cell stack 100 and sensor 104 is activated. During operation of stack 100, at step 124, sensor 104 monitors at least one operational characteristic that correlates to $dV/d(OS)$. For example, for a fuel cell operating at a constant current density, the operational characteristic measured by sensor 104 may be fuel cell voltage output or an operational characteristic which correlates to oxidant stoichiometry, because if one of these operational characteristics is known, $dV/d(OS)$ may be determined by referring to a plot of cell voltage as a function of oxidant stoichiometry (for example, FIG. 2a). Sensor 104 may measure the voltage output of fuel cell stack 100 or the voltage output of selected individual fuel cells 101, to determine $dV/d(OS)$ for stack 100 or individual fuel cells 101, respectively. Alternatively, sensor 104 may measure the concentration of oxygen in a cathode exhaust stream to determine oxidant stoichiometry, which correlates to $dV/d(OS)$. The oxygen concentration in the oxidant supply stream is known when it is pure oxygen (that is, 100%), or air (that is, about 20%). If the concentration of oxygen in a dilute oxidant supply stream is not constant, then an oxygen sensor may be used to measure the oxygen concentration upstream of fuel cell stack 100. Alternatively, if fuel cell current output and oxygen concentration in the cathode exhaust are known, controller 105 may calculate oxidant stoichiometry by determining the amount of oxygen consumed by the fuel cell to generate the

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5 electrical current.

 In another embodiment, when sensor 104
measures hydrogen concentration in the cathode
exhaust stream, the detection of hydrogen above a
threshold amount, (for example, 20 ppm), may
10 indicate that a significant amount of oxidant
starvation is occurring at a fuel cell cathode.
In this embodiment, the concentration of hydrogen
determines the severity of the oxidant starvation
that correlates to an oxidant stoichiometry for
15 the fuel cell. Accordingly, sensor 104 outputs a
signal that is representative of the measured
value of any operational characteristic that
itself, or in combination with other factors,
correlates to $dV/d(OS)$. The output signal from
20 sensor 104 may thus be received and processed by
controller 105 to calculate or infer $dV/d(OS)$ so
that controller 105 may control the oxidant
stoichiometry to reduce parasitic power
consumption.

25 In steps 126 and 128, controller 105
determines whether the operational characteristic
correlates to a $dV/d(OS)$ that is within the
desired range. At step 126, if the operational
characteristic correlates to a $dV/d(OS)$ that is
30 lower than a first predetermined value (that is,
the upper limit of the desired oxidant
stoichiometry range), then, at step 132 controller
105 causes the oxidant stoichiometry to be
decreased. At step 128, if the operational
35 characteristic correlates to a $dV/d(OS)$ that is
greater than a second predetermined value
(that is, the lower limit of the desired

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5 oxidant stoichiometry range), then, at step 134
controller 105 causes the oxidant stoichiometry to
be increased. If it is determined that the
present oxidant stoichiometry is within the
desired range (that is, the answer is "no" to the
10 questions posed in both steps 126 and 128), then
no steps are taken to change oxidant
stoichiometry. After steps 128, 132 or 134, it is
determined whether sensor 104 is still activated.

As long as sensor 104 remains activated, the
15 method repeats by returning to step 124. If
sensor 104 is no longer activated, the process
stops at step 138.

The method of FIG. 3b is substantially the
same as the method of FIG. 3a, with the addition
20 of step 125, which provides for specifically
checking whether or not oxidant starvation is
detected. For the steps that are common to the
methods of both FIG. 3b and FIG. 3a the same
reference numerals are employed.

25 The desired operating range for the oxidant
stoichiometry may have a lower limit that normally
prevents significant oxidant starvation at the
cathodes. However, localized oxidant starvation
may occur, even when oxidant stoichiometry is much
30 higher than one (for example, even when oxidant
stoichiometry is within the desired operating
range). Localized oxidant starvation may occur at
portions of the cathode where the oxidant is
prevented from accessing the catalyst, for example
35 where access is prevented by the accumulation of
water. Localized starvation conditions may result
in reduced performance, lower efficiency, and

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5 undesirable hydrogen production at the cathode.
Localized oxidant starvation is difficult to
detect because fuel cell stack 100, as a whole,
may have a positive voltage within the desired
operating range, while only a portion of the
10 cathode is oxidant starved. Efficiency may be
improved by counter-acting localized starvation,
for example, by temporarily increasing oxidant
mass flow rate through the cathode to disperse
accumulated water from the cathode.

15 Localized oxidant starvation may be detected
by monitoring for irregularities that may be
indicators of oxidant starvation, such as, for
example, detecting a threshold hydrogen gas
concentration (for example, greater than 20 ppm)
20 in the cathode exhaust stream. The actual
threshold hydrogen gas concentration selected for
a particular fuel cell or fuel cell stack will
depend upon the particular characteristics such
as, for example, the number of fuel cells in a
25 stack, the mass flow rate of oxidant stream, the
type of electrolyte, and so forth. Localized
oxidant starvation may not have a determinative
effect on the voltage output of fuel cell stack
100, but any degree of oxidant starvation may
30 result in the production of hydrogen at the
cathode. Another method of detecting a
possibility of localized oxidant starvation in
fuel cell stack 100 is monitoring the voltage
output of individual fuel cells 101. If an
35 individual fuel cell has a lower voltage output
than the other fuel cells in stack 100, this is an
indication that there may be a localized oxidant

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5 starvation problem (for example, caused by the accumulation of water in the cathode of the fuel cell with the low voltage output).

 According to the method of FIG. 3b, at step 125, if oxidant starvation is detected, then
10 oxidant stoichiometry is increased at step 134. At step 125, if oxidant starvation is not detected, then the process beginning at step 126 is essentially the same as in the process of FIG. 3a.

15 FIG. 4 illustrates another method of reducing oxidant stoichiometry to reduce parasitic power consumption wherein sensor 104 monitors for an operational characteristic that is indicative of oxidant starvation at the cathode. The method of
20 FIG. 4 starts at step 140. At step 142 an oxidant stream is supplied to a fuel cell cathode, a fuel stream is supplied to a fuel cell anode and sensor 104 is activated. At step 144 activated sensor 104 monitors for oxidant starvation at the
25 cathode. Sensor 104 sends a signal to controller 105 that indicates when oxidant starvation is detected at the cathode. Sensor 104 may detect, for example, the voltage output of fuel cell stack 100 or selected individual fuel cells within stack
30 100. Alternatively, sensor 104 may detect the concentration of oxygen or hydrogen in a cathode exhaust stream. For example, if no oxygen, or only a "very low concentration", is detected, this is a good indication that oxidant starvation may
35 be occurring at the cathode. The definition of what constitutes a "very low concentration" of oxygen depends upon the operating conditions and

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5 the properties of the reactants. For example, if
the inlet oxidant stream comprises about 20%
oxygen, then a "very low concentration" may be 5%
oxygen in the cathode exhaust stream (that is,
corresponding to an overall oxidant stoichiometry
10 of about 1.33). However, if, for example, the
inlet oxidant stream comprises 30% oxygen, then a
"very low concentration" may be 7% (that is,
corresponding to an overall oxidant stoichiometry
of about 1.30). Preferably, the concentration of
15 oxygen selected as a threshold value corresponds
to a predetermined oxidant stoichiometry, with
consideration given to the oxygen concentration in
the inlet oxidant stream. Similarly, if sensor
104 measures a hydrogen concentration above a
20 threshold amount (for example, above 20 ppm), then
it is likely that oxidant starvation is occurring
at a fuel cell cathode. The values for the
threshold concentrations depend upon the
particular characteristics of the fuel cell
25 system, such as, for example, the oxidant stream
flow rate, the size of the fuel cells and the
number of fuel cells in a stack.

At step 146, controller processes the signal
from sensor 104 and determines whether oxidant
30 starvation, or a likelihood of oxidant starvation,
has been detected. In either case, if such a
condition is detected, then the oxidant
stoichiometry is increased at step 148. If, at
step 146 it is determined that there is no
35 indication of oxidant starvation at the fuel cell
cathodes, then the oxidant stoichiometry is
decreased at step 150.

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5 After either step 148 or 150, controller 105
checks at step 152 whether sensor 104 is still
activated. If sensor 104 is no longer activated,
then the method stops at step 154. As long as
sensor 104 remains activated, the method of
10 controlling the oxidant stoichiometry repeats by
returning to step 144.

 The logic diagrams of FIGs. 5-12 are directed
to examples of methods wherein the operational
characteristic measured by sensor 104 is the
15 concentration of hydrogen gas in a cathode exhaust
stream. In view of the present disclosure,
persons skilled in the art will understand that,
alternative operational characteristics such as,
for example, fuel cell voltage output or oxidant
20 concentration, may also be used in conjunction
with the methods set out in the logic diagrams of
FIGs. 5-12. That is, sensor 104 may be a sensor
that measures any operational characteristic of
the fuel cell system that may be employed by
25 controller 105 to determine $dV/d(OS)$ and/or the
presence of oxidant starvation at the cathode.

 FIG. 5 is a logic diagram for a method
wherein controller 105 determines whether to
increase or decrease oxidant stoichiometry based
30 upon whether the hydrogen gas concentration
measured by sensor 104 exceeds a predetermined
threshold concentration (TC). The method starts
at step 160. At step 162 sensor 104 monitors the
cathode exhaust stream for hydrogen gas. At step
35 164 controller 105 determines whether the hydrogen
gas concentration measured at step 162 exceeds TC.
Minute quantities of hydrogen in the cathode

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exhaust stream may not be indicative of a problem. Accordingly, TC is a threshold hydrogen gas concentration that is empirically known for a particular fuel cell to be indicative of oxidant
5 starvation at the cathodes. For example, for a Ballard® MK V fuel cell, controller 105 may be programmed so that TC is a hydrogen concentration of 20 ppm. Until sensor 104 detects a hydrogen concentration greater than or equal to TC,
10 controller 105 reduces oxidant stoichiometry at step 166, typically by reducing the mass flow rate of the oxidant supply stream (thereby reducing parasitic power consumption).

When sensor 104 detects a hydrogen
15 concentration greater than TC, this is an indication that oxidant starvation is occurring (or is likely to occur) and controller 105 proceeds to step 168. At step 168, controller 105 determines whether the oxidant stream mass flow
20 rate is greater than or equal to a desired maximum. If the oxidant stream mass flow rate is not greater than or equal to the desired maximum, then oxidant stoichiometry is increased at step 170. Normally oxidant stoichiometry is increased
25 by controlling motor 112 to increase the speed of mechanical device 111. Alternatively, other measures may be used instead or in combination with controlling motor 112 to increase oxidant stoichiometry. For example, oxidant stoichiometry
30 may be increased by reducing electrical power output without a corresponding reduction in oxidant stream mass flow rate.

If, however, it is determined at step 168

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5 that the oxidant stream mass flow rate is in fact
greater than or equal to the desired maximum mass
flow rate, then, controller 105 proceeds to step
172 and determines whether to generate a warning
signal (step 174) while continuing to operate the
10 fuel cell system, or, shut down (cease operation
of) fuel cell stack 100 at step 176. At step 172,
controller 105 makes its determination by
considering whether the hydrogen gas concentration
is higher or lower than a predetermined
15 concentration limit (CL). CL is typically a value
much greater than TC. If the hydrogen
concentration is higher than CL, this indicates
that a much higher than normal concentration of
hydrogen is present at the cathode. For example,
20 if the fuel is hydrogen, a significant amount of
fluid leakage between the anode and cathode may
cause the hydrogen concentration in the cathode
exhaust stream to exceed CL, and such a condition
warrants shutting the fuel cell down so that the
25 cause of the elevated hydrogen concentration may
be investigated.

 If the fuel cell system comprises an array of
fuel cell stacks with each stack monitored in the
manner depicted in FIG. 5, then the array may
30 continue to produce electrical power, but with one
stack shut down.

 As noted above, operational characteristics
other than hydrogen concentration in the cathode
exhaust stream may be employed instead. For
35 example, with reference to the logic diagram of
FIG. 5, alternate equivalent steps are described
herein for a sensor that measures fuel cell

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5 voltage output. Alternate steps are described
using the same reference numerals, but with a
suffix "a". According to the alternate method,
after starting the method at step 160a, a fuel
cell voltage output sensor monitors output voltage
10 at step 162a. At step 164a, controller 105
determines if the output voltage is lower than a
predetermined threshold output voltage (TOV). In
this example, the predetermined TOV preferably
corresponds to a condition when there is a
15 potential for oxidant starvation at the fuel cell
cathode. If voltage output is lower than TOV, the
oxidant stoichiometry is increased at step 170a,
but not before first checking, at step 168a, that
the oxidant stream mass flow rate is not already
20 greater than or equal to a desired maximum. If
fuel cell voltage output is not less than TOV,
then controller 105 decreases oxidant
stoichiometry at step 166a. At step 168a, if it
is determined that oxidant stream mass flow rate
25 is greater than or equal to a desired maximum,
then controller 105 proceeds to step 172a. At
step 172a, if the sensor measures an output
voltage that indicates cell reversal, then
controller 105 shuts down the fuel cell at step
30 176a. If cell reversal is not detected, then the
fuel cell system may continue to operate, but with
controller 105 generating a warning signal at step
174a.

FIG. 6 is a logic diagram which illustrates a
35 method wherein controller 105 determines an
appropriate action with reference to a
predetermined desired oxidant stoichiometry for

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5 the instant electrical power output. The method starts at step 180 with the supply of reactants to fuel cell stack 100 and activation of sensor 104.

At step 182, sensor 104 monitors the cathode exhaust stream for hydrogen gas. At step 184, if
10 sensor 104 measures a hydrogen gas concentration less than threshold concentration TC, the controller 105 decreases the oxidant stoichiometry at step 186. If sensor 104 does not detect a hydrogen gas concentration less than TC, then, at
15 step 188, controller 105 refers to a look up table to determine the desired oxidant mass flow rate for the instant electrical power output. At step 190, controller 105 determines whether the actual oxidant mass flow rate is more than a
20 predetermined amount (for example, P%) higher than the desired oxidant mass flow rate. If the actual oxidant mass flow rate is not already greater than P% higher than the desired oxidant mass flow rate, then controller 105 increases the oxidant
25 stoichiometry at step 192. If, however, the actual oxidant mass flow rate is greater than P% higher than the desired oxidant mass flow rate, then, at step 194, controller 105 generates a warning signal or shuts down the fuel cell. If
30 controller 105 generates a warning signal, it may also control the fuel cell system to reduce the electrical power output or limit the peak power output. As with the embodiment of FIG. 5, in the method of FIG. 6, the magnitude of the measured
35 hydrogen gas concentration may be used to determine the appropriate action (that is, a warning signal or shutting down (ceasing operation

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5 of) fuel cell stack 100).

In the logic diagrams of FIGs. 7-9, the methods comprise determining whether the hydrogen gas concentration detected in the cathode exhaust stream is increasing or decreasing, and using this
10 information to determine what action is appropriate in response to the changing hydrogen gas concentration.

With reference to FIG. 7, the method starts at step 200, by supplying reactants to fuel cell
15 stack 100 and activating hydrogen sensor 104. At step 202 sensor 104 begins monitoring the cathode exhaust stream for hydrogen gas. Once activated, sensor 104 measures the hydrogen gas concentration in the cathode exhaust stream. Sensor 104
20 measures the instant hydrogen gas concentration (H) and controller 105 calculates dH/dt where dH is the change in H and dt is the change in time (at constant time intervals). By calculating whether dH/dt is positive, negative, or zero,
25 controller 105 determines whether H is increasing, decreasing, or constant, respectively.

Step 204 follows step 202. At step 204, controller 105 determines whether the instant H is higher than a predetermined threshold
30 concentration (TC). If not, then controller determines that there is no oxidant starvation at the cathode and there is excess oxygen at the cathode. Accordingly, at step 206, controller 105 decreases the oxidant stoichiometry, for example,
35 by reducing the speed of motor 112 to decrease the oxidant stream mass flow rate supplied to fuel cell stack 100. Oxidant stoichiometry is thus

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5 advantageously controlled to reduce parasitic power consumption.

However, if at step 204, controller determines that the instant H is greater than TC , controller 105 advances to step 208 to determine
10 whether dH is negative (that is, whether hydrogen concentration is decreasing). If controller 105 determines that dH/dt is negative, controller 105 preferably returns to step 202 without taking any positive action to counteract
15 the detection of a hydrogen gas concentration greater than TC . However, at step 208, if controller 105 determines that dH/dt is not negative (that is, H is greater than TC and the hydrogen concentration is either constant or
20 increasing), then, controller 105 advances to step 210 and determines whether the oxidant stream mass flow rate is greater than or equal to a desired maximum mass flow rate. If oxidant stream mass flow rate is not greater than or equal to the
25 desired mass flow rate, then controller 105 proceeds to step 212 and increases the oxidant stoichiometry, preferably by a predetermined increment. Controller 105 then returns to step 202 to determine the effect of the increase in
30 oxidant stoichiometry and to repeat the method.

At step 210, if controller 105 determines that the oxidant stream mass flow rate is already greater than or equal to a desired maximum mass flow rate, then controller 105 advances to step
35 214 and generates a warning signal and may eventually shut down fuel cell stack 100 (step 220). In the embodiment shown in FIG. 7,

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5 controller 105 may initiate additional steps 216
and 218 before proceeding to shut down fuel cell
stack 100. Since the oxidant stream mass flow
rate is already greater than or equal to the
desired maximum, controller 105 proceeds to step
10 214 to generate a warning signal and then to step
216 to attempt a corrective action other than
increasing the oxidant stream mass flow rate. One
or more leaks within the fuel cell may be the
cause for detecting an excessive amount of
15 hydrogen gas in the cathode exhaust stream. For
example, hydrogen could be leaking from the anode
fluid passages to the cathode fluid passages, or a
leak in the oxidant delivery subsystem could
prevent a sufficient supply of oxygen from being
20 directed to the cathode.

At step 216, controller 105 reduces the fuel
pressure in fuel cell stack 100. If one or more
leaks between the anode and the cathode are the
cause for detecting hydrogen gas at the cathode,
25 then reducing fuel pressure at the anode may
reduce the rate of transfer of fuel from the
anodes to the cathodes. Fuel pressure may be
adjusted, for example by adjusting a pressure
control valve or reducing the speed of a fuel
30 compressor. Since the fuel stoichiometry may
initially be greater than one, the reduction in
fuel pressure may not have an immediate effect on
electrical power output.

At step 218 controller 105 determines whether
35 electrical power output is in fact less than
electrical power demand. If electrical power
output from fuel cell stack 100 is less than

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electrical power demand, then controller 105 proceeds to shut down fuel cell stack 100. If, however, electrical power output continues to match electrical power demand, then controller 105
5 returns step 202 and the fuel cell system continues to operate while the method is repeated. Meanwhile, the warning signal generated at step 214 alerts the operator that there is a problem that needs to be investigated to determine why H
10 is greater than TC.

In FIG. 8, the method starts at step 230 and advances immediately to step 232 where sensor 104 begins to monitor the cathode exhaust stream to measure hydrogen gas concentration (H), and
15 controller 105 calculates dH/dt . At step 234, controller 105 determines whether H is greater than threshold concentration (TC) or, at step 238, whether dH/dt is negative. The logic of steps 232, 234, 236, and 238 is substantially the same
20 as the logic of corresponding steps shown in FIG. 7 (that is, 202, 204, 206, and 208, respectively).

However, in the method of FIG. 8, if it is determined at step 238 that dH/dt is not negative, then controller 105 proceeds to step 240 and
25 considers whether the fuel cell voltage is less than a predetermined voltage V_0 . For example, in a preferred embodiment, voltage V_0 is assigned a value between zero and about 200 mV. The selected value for voltage V_0 is preferably greater than
30 zero because localized oxidant starvation may produce hydrogen even though the overall cell voltage is still positive. For example, a hydrogen concentration of about 20 ppm may be

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5 measured in the cathode exhaust stream when fuel
cell stack 100 has an average cell voltage of
about 100 mV. Accordingly, if controller 105
determines that there is a combination of a
positive dH/dt , and a cell voltage less than V_o ,
10 the cause of these two conditions may be oxidant
starvation. However, if the cell voltage is
greater than V_o , this is an indication that cell
reversal is probably not the cause for detecting
hydrogen gas in the cathode exhaust.

15 Therefore, if controller 105 determines at
step 240 that cell voltage is not less than V_o ,
controller 105 determines that oxidant starvation
is not likely the cause for H being greater than
TC (that is, because oxidant starvation would be
20 accompanied by a significantly reduced fuel cell
voltage). More likely, the cause of H being
greater than TC is one or more fluid leaks between
the anode and cathode. Accordingly, controller
105 proceeds to step 246 and reduces the fuel
25 pressure in fuel cell stack 100 to confirm that
fluid leaks are indeed the reason for H being
greater than TC. Steps 248 and 250 are
essentially the same as steps 218 and 220 in the
method of FIG. 7.

30 In order to better detect localized oxidant
starvation conditions in fuel cell stack 100, the
voltage of the individual fuel cells may be
monitored. Alternatively, a more simplified cell
voltage monitoring system may be employed to
35 measure the cell voltage of selected fuel cells in
a stack, or the average cell voltage of more than
one fuel cell in a fuel cell stack. Accordingly,

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5 in the embodiments of FIG. 8 and 9, the measured voltage V_o may be an individual fuel cell voltage, an average voltage of a plurality of fuel cells, or the average voltage of all the fuel cells in fuel cell stack 100.

10 In the embodiment of FIG. 8, at step 240, if the fuel cell voltage is less than V_o , controller 105 proceeds to step 242 and checks whether the oxidant stream mass flow rate is already greater than or equal to the desired maximum. If yes,
15 this is an indication that oxidant starvation is not likely the reason for H being greater than TC, and controller proceeds to step 246. Steps 246, 248 and 250 are similar to corresponding steps 216, 218, and 220 of FIG. 7. If the oxidant
20 stream mass flow rate is less than the desired maximum, at step 244 controller 105 increases the oxidant stoichiometry, for example, by causing motor 112 to speed up so mechanical device 111 increases the oxidant stream mass flow rate to
25 fuel cell stack 100.

The method illustrated by FIG. 9 has steps 260, 262, 264, 266, 268, 270, 272, and 274 in common with respective steps 230, 232, 234, 236, 238, 240, 242, and 244 of the method shown in FIG.
30 8. In FIG. 9, steps 276 and 278 are substituted for steps 246 and 248 in FIG. 8, and additional steps 280 and 284 are added following step 278. If at step 270, controller 105 determines that the fuel cell voltage is not less than V_o , controller
35 105 may proceed to step 276 to take further action to confirm that a leak is the likely source of the hydrogen gas in the cathode exhaust stream. At

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step 276, the pressure of the oxidant and fuel stream are adjusted to regulate the pressure differential between the oxidant and fuel stream fluid passages. For example, controller 105 may
5 cause an increase in the pressure differential between the oxidant and fuel passages to determine whether this has a corresponding effect on H detected by sensor 104. If, at step 278, a corresponding effect is detected (that is, the
10 adjustment of the pressure differential in step 276 influenced the hydrogen gas concentration in the cathode exhaust stream), then it is determined that a leak is the likely source of the hydrogen gas and controller 105 generates a warning signal
15 at step 280. Then controller 105 selects whether to continue operating at a reduce power output (step 284) or to shut down the fuel cell (step 282) If controller 105 selects step 284, it may further regulate the oxidant and fuel pressures to
20 balance the reactant pressures so that the pressure differential is approximately zero to thereby reduce the amount of hydrogen which is transferred from the fuel passages to the oxidant passages. The choice between step 282 and step 284
25 may be made based on the variance between the electrical power output and the electrical power demand. For example, if the electrical power output is more than a predetermined amount less than electrical power demand, step 282 is selected
30 and fuel cell stack 100 is shut down.

Controller 105 may also proceed to step 276 from step 272, if the fuel cell voltage is less than V_0 and the oxidant mass flow rate is greater

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5 than or equal to a desired maximum.

FIGs. 10 and 11 depict preferred embodiments of a method of controlling the oxidant stoichiometry wherein sensor 104 is periodically activated to detect the presence of hydrogen gas in the cathode exhaust stream. The delay between periodic activation of sensor 104 may vary depending upon several variables. A short delay may be employed for applications where the electrical power output is continually changing, for example where fuel cell stack 100 is supplying electrical power to a drive motor for a vehicle. In such applications, the delay may only be long enough to allow the effects of any corrective actions to be determined. A short delay is desired so that sensor 104 is activated with sufficient frequency to be responsive to changes in electrical power output. Longer delays may be employed by applications such as stationary power plants, which tend to operate to produce a more constant electrical power output.

With reference to FIG. 10, the control logic starts by activating controller 105 at step 290. At step 292, the initial value for the hydrogen gas concentration (C_0) is set to zero. At step 294 controller 105 sets counter number "n" to one. At step 296, controller 105 activates sensor 104 to detect the presence of hydrogen gas in a cathode exhaust stream. Sensor 104 emits an output signal that is representative of the hydrogen gas concentration (C_n). This output signal is sent to controller 105, which determines at step 298 whether C_n is greater than the

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5 threshold concentration TC. If C_n is not greater
than TC, then it is assumed that there are no
problems with leaks or oxidant starvation. To
improve efficiency by reducing parasitic
electrical loads, controller 105 proceeds to step
10 300 and decreases the oxidant stoichiometry by a
predetermined amount (for example, a fixed
increment or a percentage). Then at step 302,
controller 105 increases the value of n by one
before proceeding to step 304. At step 304
15 controller 105 waits for a predetermined delay
period to elapse before returning to step 296 to
re-activate sensor 104. Therefore, so long as
there is no oxidant starvation, and no leaks or
other sources of hydrogen gas in the cathode
20 exhaust, controller 105 will continue to loop
through steps 296, 298, 300, 302 and 304. In this
way, the oxidant stoichiometry is reduced to about
one or until C_n is greater than TC. Decreasing
oxidant stoichiometry in this manner reduces
25 parasitic power consumption.

At step 298, when C_n is greater than TC,
controller 105 proceeds to step 306 and compares
 C_n to the previously measured hydrogen gas
concentration, $C_{(n-1)}$. If $C_n - C_{(n-1)}$ is negative,
30 this indicates that the hydrogen gas concentration
has decreased and controller 105 returns to step
296 via steps 302 and 304. As long as the
hydrogen gas concentration is decreasing,
controller 105 does not actively take any
35 corrective action. At step 302, counter number n
is increased by one. Step 304 is the delay step.

However, if at step 306, controller 105

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5 determines that $C_n - C_{(n-1)}$ is positive, this indicates that the hydrogen gas concentration has increased from the previous measurement. This prompts controller 105 to take corrective action by proceeding to step 308. At step 308,
10 controller 105 determines from the present oxidant stream mass flow rate whether or not the oxidant stream mass flow rate is greater than or equal to the desired maximum. If not, controller 105 proceeds to step 310 and increases the oxidant
15 stoichiometry, for example, by causing motor 112 to speed up so mechanical device 111 increases the mass flow rate of the oxidant stream supplied to fuel cell stack 100. After step 310, controller 105 eventually loops back to step 296 after
20 performing intermediate steps 302 (increasing counter number n by one) and 304 (the delay step). This corrective action presumes that the cause of the increased concentration of hydrogen gas in the cathode exhaust is oxidant starvation at the
25 cathode.

However, if oxidant starvation is not the cause for the increasing hydrogen gas concentration in the cathode exhaust, the oxidant stream mass flow rate will soon be increased to
30 the maximum mass flow rate. Then, at step 308, controller 105 will recognize that oxidant starvation is not the cause of the problem and controller 105 will proceed to step 312.

At step 312, controller 105 reduces the
35 pressure in the fuel stream by a predetermined amount (for example, a fixed increment or a percentage). If the source of the hydrogen gas

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5 detected in the cathode exhaust is a leak in fuel
cell stack 100, reducing the pressure in the fuel
stream may have a corresponding effect on the
leakage rate. Since the fuel stoichiometry may
initially be greater than 1.0 (for example, a fuel
10 stoichiometry of 1.5 or 2.0 is common), decreasing
the fuel mass flow rate may not initially affect
the electrical power output. At step 314,
controller 105 checks whether the electrical power
output is less than the electrical power demand.
15 If the electrical power output is not less than
the electrical power demand, controller 105 loops
back to step 296 after performing intermediate
steps 302 and 304. However, if the electrical
power output is less than electrical power demand,
20 this indicates that there is a problem with fuel
cell stack 100 that prevents it from performing at
this capacity. In this case, controller 105
proceeds to step 316 where controller 105 either
generates a warning signal or causes fuel cell
25 stack 100 to shut down.

With regard to steps 300, 310, and 312 of
FIG. 10, the respective oxidant stoichiometry or
fuel pressure may be increased and decreased by a
predetermined fixed percentage or a fixed
30 increment. For example, the fixed percentage
change in oxidant stoichiometry or fuel pressure
may be 1 or 2%.

With reference to FIG. 11, steps 320, 322,
324, 326, 328, 332, 334, 336, 340, 342, 344, and
35 346 are substantially the same as steps 290, 292,
294, 296, 298, 302, 304, 306, 308, 312, 314 and
316 of FIG. 10, respectively. However, in the

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5 method of FIG. 11, controller 105 attempts to maintain an oxidant stoichiometry that is a desired amount above 1.0 (that is, $1.0 + Y$). For example Y may be, for example, 0.05, 0.10, 0.20, or even about 0.50 (that is, about 50%).

10 Preferably Y represents an increase of less than 50% to avoid excessive parasitic power consumption. In this way, a small surplus oxidant buffer is provided to reduce the occurrence of oxidant starvation conditions that might produce

15 hydrogen gas in the cathode exhaust. The oxidant stoichiometry is still preferably much less than the oxidant stoichiometries used in conventionally operated fuel cells so there is still a substantial reduction in the parasitic electrical

20 load caused by the operation of oxidant stream mechanical device 111.

In the embodiment of FIG. 11, after step 328, if controller 105 determines that the hydrogen gas concentration (C_n) is less than threshold

25 concentration TC, then controller loops back to step 326 via steps 330, 332 (increasing counter number n by one), and 334 (the delay step). At step 330 controller 105 decreases the oxidant stoichiometry by a predetermined increment (X),

30 where X may be, for example, about 0.1 or about 0.2. Accordingly, if, for example, X is 0.1 and oxidant stoichiometry is 1.4, at step 330 controller 105 would adjust oxidant stream mass flow rate to decrease oxidant stoichiometry to

35 1.3. Thus, while there is an excess of oxidant at the cathode, oxidant stream mass flow rate is reduced through the loop comprising steps 326,

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5 328, 330, 332 and 334 until controller 105
determines at step 328 that C_n is greater than TC,
whereupon controller 105 proceeds to step 336.

10 Step 336 of FIG. 11 performs substantially
the same function as step 306 of FIG. 10. That is,
if the hydrogen gas concentration is decreasing
(that is, $C_n - C_{(n-1)}$ is negative), then controller
returns to step 326 via steps 332 and 334.

Controller 105 continues to periodically monitor
 C_n without taking any action to actively increase
15 oxidant stoichiometry. However, if controller 105
determines at step 328 that $C_n > TC$, and at step
336, that the hydrogen gas concentration is
constant or increasing (that is, $C_n - C_{(n-1)}$ is not
negative), the controller 105 proceeds to step
20 338.

At step 338, if controller 105 determines
that the previous measurement of the hydrogen gas
concentration was less than TC, controller 105
increases the oxidant stream mass flow rate by (X
25 + Y) at step 348. For example, when TC
corresponds to an oxidant stoichiometry of about
one, in the method of FIG. 11, at step 348 oxidant
stoichiometry is controlled so that it is
approximately $1.0 + Y$. That is, in the previous
30 loop, hydrogen concentration $C_{(n-1)}$ was less than
TC, but decreasing oxidant stoichiometry by X
caused the hydrogen concentration in the next loop
(that is, the present loop) to be higher than TC.

Accordingly, in the previous loop, $C_{(n-1)}$ was close
35 in value to TC, and since in this example, a
measured hydrogen concentration of about TC
corresponds to an oxidant stoichiometry of about

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5 one, increasing oxidant stoichiometry by $(X + Y)$,
results in an oxidant stoichiometry of about $1.0 + Y$.

If at step 338, controller 105 determines
that the previously measured hydrogen gas
10 concentration was greater than TC, controller 105
proceeds to step 340 to determine whether the
oxidant stream mass flow rate is already greater
than or equal to a desired maximum. If not,
controller 105 proceeds to step 350 and increases
15 the oxidant stoichiometry by Z. The value of Z is
preferably greater than $(X + Y)$ so that when C_n is
higher than TC for more than one loop, the oxidant
stoichiometry is increased more rapidly.

However, if at step 340 controller 105
20 determines that oxidant stream mass flow rate is
already greater than or equal to a desired
maximum, controller 105 proceeds to step 342.
Steps 342, 344 and 346 are substantially the same
as respective steps 312, 314 and 316 in FIG. 10.

25 FIG. 12 depicts another preferred embodiment
of a method of controlling the supply of oxidant
to a fuel cell. In the method of FIG. 12 the
preferred oxidant stream supply mass flow rate for
a predetermined oxidant stoichiometry is
30 calibrated for a range of specific electrical
power outputs. During a calibration procedure,
controller 105 determines the desired oxidant
stream mass flow rates (for a particular oxidant
stream composition) for selected electrical power
35 outputs and stores the desired mass flow rates in
a look-up table. When this method is employed,
the electrical power demand determines the oxidant

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5 stream mass flow rate. That is, a controller
monitors the electrical power demand and sets the
oxidant mass flow rate with reference to a look-up
table, which indicates the desired oxidant stream
mass flow rate for selected electrical power
10 demands. In this way, the oxidant stoichiometry
is controlled to reduce the amount of excess
oxidant supplied to fuel cell stack 100, thereby
reducing parasitic electrical power consumption.

FIG. 12 depicts a calibration procedure for
15 determining the oxidant stream mass flow rate for
different selected electrical power outputs. When
fuel cell stack 100 is actually operating, if the
electrical power output is between selected loads
in the look-up table, the desired oxidant stream
20 mass flow rate may be determined by interpolating
between selected electrical loads to determine the
desired oxidant stream mass flow rate.

In an alternative embodiment, the calibration
procedure may be used to calibrate the oxidant
25 supply system directly. For example, in a system
that employs an oxidant compressor, the system is
calibrated to control the speed of the compressor
in response to selected electrical power demands.
In this way, oxidant stream mass flow rate need
30 not be measured and with each calibration, the
calibration process will automatically compensate
for any degradation in compressor performance over
its operational lifetime.

The calibration procedure depicted in FIG. 12
35 may be executed periodically when fuel cell stack
100 is being serviced for regular maintenance.
The procedure begins at step 360, by activating

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5 controller 105. In this embodiment, because calibration is typically done during maintenance periods, controller 105 and sensor 104 may be detachable from the fuel cell system. In this way, the same equipment may be used to calibrate a
10 plurality of fuel cell stacks. A coupling may be provided so that a sensing element may be inserted into the fuel cell so that the sensing element is exposed to the cathode exhaust stream.

At step 362 an electrical power output is
15 selected and fuel cell stack 100 is operated to produce that electrical power output. The selected electrical power output may be any electrical power output within the operating range of fuel cell stack 100. A plurality of electrical
20 power outputs are typically selected during the calibration procedure so it is convenient to start with an electrical power output at the low end of the range; progressively higher electrical power outputs may be subsequently selected to complete
25 the calibration procedure.

At step 364, controller 105 accesses a look-up table to determine the previously calibrated oxidant stream mass flow rate for the selected electrical power output. Controller 105 then sets
30 the oxidant stream mass flow rate so that it is initially 1% higher than the previously calibrated oxidant stream mass flow rate for the selected electrical power output. It is preferable to calibrate the oxidant stream mass flow rate by
35 starting with a surplus oxidant stream mass flow rate rather than a shortage of oxidant, since this precaution avoids initiating the calibration

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5 procedure in an oxidant starvation mode.

At step 366 sensor 104 is activated to detect the presence of hydrogen gas in the cathode exhaust stream. If fuel cell stack 100 is initially supplied with a surplus of oxidant, at
10 step 368, the hydrogen gas concentration C is expected to be less than a threshold concentration TC. If controller 105 determines that C is not greater than TC, controller 105 continues to decrease the oxidant stream mass flow rate by
15 looping back to step 366 via step 370. With each loop through step 370, controller 105 decreases the oxidant stream mass flow rate by increment A. For example, the value of increment A may correspond to stoichiometry reductions, of say
20 0.05 or 0.1 so that oxidant stoichiometry is reduced by that amount each time step 370 is performed. Because the selected load is constant during the calibration procedure, changes in the oxidant stream mass flow rate result in
25 corresponding changes in the oxidant stoichiometry. The accuracy of the calibration procedure may be increased by decreasing the value of A so that more calibration loops are performed using smaller incremental reductions in the
30 oxidant stream mass flow rate.

When the oxidant stream mass flow rate is finally reduced so that hydrogen gas concentration C is greater than TC, controller 105 proceeds to step 372 and increases the oxidant stream mass
35 flow rate by increment B. Next, at step 374, controller 105 resets the look-up table value so that the desired oxidant stream mass flow rate in

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the look-up table matches the current mass flow rate for the selected electrical power output. Finally, controller 105 may choose to stop the calibration procedure at step 378, or to select a
5 different electrical power output for oxidant stream mass flow rate calibration at step 376. If step 376 is chosen, a different electrical power output is selected and controller 105 returns to step 364 where the calibration procedure begins
10 again for the newly selected electrical power output.

The value of A and B may be the same or B may be higher than A. When B is higher than A, the look up table values will be calibrated so that
15 there will be a surplus of oxidant supplied to the cathode. The greater the difference between the values for B and A, the greater will be the surplus. A surplus supply of oxidant helps to reduce the likelihood of causing oxidant
20 starvation conditions that might produce hydrogen gas in the oxidant stream. When the power output is expected to be dynamic, a higher oxidant stoichiometry may be preferred to prevent oxidant starvation during transitional periods when
25 electrical power output is changing. The value of B is preferably selected so that the oxidant stoichiometry is generally less than two and preferably between about one and about 1.5.

30

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5 in the art without departing from the spirit and scope of the present disclosure, particularly in light of the foregoing teachings.

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WHAT IS CLAIMED IS:

1. A method of operating a fuel cell system, said system comprising a fuel cell power
5 generating subsystem having at least one fuel cell, and an oxidant delivery subsystem, comprising a mechanical device for supplying an oxidant stream to a cathode of said fuel cell, wherein said method comprises controlling said
10 mechanical device to reduce parasitic power consumption by reducing oxidant stoichiometry until $dV/d(\text{oxidant stoichiometry})$ is greater than a predetermined value.
- 15 2. The method of claim 1 wherein said fuel cell is a solid polymer electrolyte fuel cell.
3. The method of claim 1 wherein said predetermined value of $dV/d(\text{oxidant stoichiometry})$
20 is greater than or equal to 0.02 volt.
4. The method of claim 1 wherein said predetermined value of $dV/d(\text{oxidant stoichiometry})$ is between about 0.3 volt and about 7.0 volts.
25
5. The method of claim 1 wherein said predetermined value of $dV/d(\text{oxidant stoichiometry})$ is determined when cell voltage decreases below a threshold voltage.
30
6. The method of claim 1 wherein said predetermined value of $dV/d(\text{oxidant stoichiometry})$

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is determined when oxidant starvation is occurring at the cathode electrochemically active area.

7. The method of claim 5 wherein the determination that $dV/d(\text{oxidant stoichiometry})$ is greater than said predetermined value is made by detecting a threshold concentration of hydrogen in the cathode exhaust stream.

8. The method of claim 1 wherein said predetermined value for $dV/d(\text{oxidant stoichiometry})$ is determined when said oxidant stoichiometry is less than a threshold value that is between about 1.0 and about 2.0.

9. The method of claim 8 further comprising measuring oxygen concentration within said cathode exhaust stream to calculate oxidant stoichiometry to determine when said oxidant stoichiometry is less than said threshold oxidant stoichiometry.

10. The method of claim 1 wherein said predetermined value for $dV/d(\text{oxidant stoichiometry})$ is determined when the value of said cell voltage is less than a threshold cell voltage and cell voltage is monitored to determine when to cease reducing oxidant stoichiometry.

11. The method of claim 10 further comprising controlling said oxidant delivery subsystem to maintain voltage output within a predetermined voltage range which corresponds to

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an oxidant stoichiometry range between about 1.0 and about 2.0, wherein said threshold cell voltage corresponds to the cell voltage at the lower limit of said predetermined voltage range.

5

12. The method of claim 1 wherein said oxidant stream is a dilute oxidant stream.

13. The method of claim 12 wherein said
10 oxidant stream is air.

14. The method of claim 8 further comprising controlling said mechanical device to maintain oxidant stoichiometry at about 1.0 during steady
15 state operating conditions.

15. The method of claim 1 wherein said mechanical device is selected from the group consisting of a compressor, a fan, a pump, and a
20 blower, and wherein oxidant stoichiometry is reduced by reducing the speed of said mechanical device.

16. The method of claim 9 wherein the method
25 of determining oxidant stoichiometry further comprises monitoring oxygen concentration in an oxidant supply stream that is directed to said cathode.

17. The method of claim 9 further comprising
30 monitoring electrical current output from said fuel cell power generating subsystem.

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18. A method of operating a fuel cell, wherein said fuel cell has a cathode supplied with an oxidant stream and an anode supplied with a fuel stream, said method comprising:

- 5 (a) monitoring a cathode exhaust stream downstream of said cathode to detect hydrogen gas concentration; and
- (b) decreasing oxidant stoichiometry when
10 said hydrogen gas concentration is less than a first threshold concentration.

19. The method of claim 18 wherein said fuel cell is one of a plurality of fuel cells arranged in a fuel cell stack and said cathode exhaust
15 stream is monitored downstream of a plurality of cathodes associated with said plurality of fuel cells.

20. The method of claim 18 further
20 comprising increasing said oxidant stoichiometry when said hydrogen gas concentration is greater than a second threshold concentration.

21. The method of claim 20 wherein said
25 first threshold concentration is the lower detection limit of a hydrogen sensor used to monitor said cathode exhaust stream.

22. The method of claim 20 wherein said
30 second threshold concentration is about 20 ppm of hydrogen.

23. The method of claim 20 wherein said
35 second threshold concentration is greater than said first threshold concentration.

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24. The method of claim 20 wherein said
oxidant stoichiometry is adjusted by adjusting the
oxidant concentration in said oxidant stream
5 supplied to said cathode.

25. The method of claim 20 wherein said
oxidant stoichiometry is adjusted by adjusting the
electrical power output of said fuel cell.
10

26. The method of claim 20 wherein said
oxidant stoichiometry is adjusted by adjusting the
mass flow rate of said oxidant stream supplied to
said cathode.
15

27. The method of claim 26 wherein said
oxidant stream mass flow rate is adjusted by
adjusting the speed of a mechanical device which
supplies said oxidant stream to said cathode.
20

28. The method of claim 26 wherein said
oxidant stream mass flow rate is adjusted by a
fixed amount or by a fixed percentage of the
instantaneous oxidant stream mass flow rate.
25

29. The method of claim 26 wherein said
oxidant stoichiometry is adjusted by adjusting
said oxidant stream mass flow rate by an amount
that is dependent upon the magnitude of the
30 detected hydrogen gas concentration.

30. The method of claim 18 further
comprising steps for reducing said hydrogen gas
concentration within said cathode exhaust stream
35 when said hydrogen gas concentration is greater

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than a second threshold concentration, wherein said steps comprise comparing said oxidant stream mass flow rate to a maximum desired mass flow rate, and

- 5 (a) if said oxidant stream mass flow rate is less than said maximum desired mass flow rate, increasing said oxidant mass flow rate; and
- 10 (b) if said oxidant mass flow rate is greater than or equal to said maximum desired mass flow rate, ceasing operation of said fuel cell if said hydrogen gas concentration is greater than a third concentration threshold
- 15 which is greater than said first and second concentration thresholds; and generating a warning signal and continuing to operate said fuel cell if said hydrogen gas concentration is less than said third
- 20 concentration threshold.

31. The method of claim 20 wherein said method further comprises taking no steps to change said oxidant stoichiometry when said hydrogen gas

25 concentration is between said first and second threshold concentrations.

32. The method of claim 26 further comprising, after increasing said oxidant mass

30 flow rate, comparing said oxidant mass flow rate to a desired oxidant mass flow rate for the instantaneous fuel cell electrical power output, and generating a warning signal or ceasing operation of said fuel cell if said oxidant mass

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flow rate is more than a predetermined amount greater than said desired oxidant mass flow rate.

33. The method of claim 30 wherein said
5 desired oxidant mass flow rate is determined from a look-up table.

34. The method of claim 18 wherein said
10 monitoring comprises continuously monitoring said cathode exhaust stream for said hydrogen gas concentration and determining whether said hydrogen gas concentration is increasing or decreasing, and when said hydrogen gas
15 concentration is greater than a second threshold concentration, said method further comprises:
maintaining a substantially constant oxidant stoichiometry when said hydrogen concentration is decreasing; and
increasing said oxidant stoichiometry
20 when said hydrogen concentration is increasing.

35. The method of claim 34 further comprising:
25 generating a warning signal if said hydrogen gas concentration is greater than said first threshold concentration, said hydrogen gas concentration within said cathode exhaust stream is increasing, and
30 said oxidant stream is flowing at a maximum desired mass flow rate.

36. The method of claim 35 further comprising controlling said fuel stream to

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decrease fuel stream pressure when said warning signal is generated.

37. The method of claim 36 further
5 comprising monitoring electrical power output of
said fuel cell, and continuing to operate said
fuel cell if said electrical power output is not
less than electrical power demand, and ceasing
operation of said fuel cell if said electrical
10 power output is a predetermined amount less than
said electrical power demand.

38. The method of claim 36 further
comprising checking electrical power output and,
15 continuing to operate said fuel cell if
said electrical power output is not less than
electrical power demand, and
generating a warning signal if said
electrical power output is less than said
20 electrical power demand and then selecting
between operating at a reduced electrical
power output and ceasing operation of said
fuel cell.

25 39. The method of claim 18 wherein said
monitoring comprises continuously monitoring said
cathode exhaust stream for said hydrogen gas
concentration and determining whether said
hydrogen gas concentration is increasing or
30 decreasing, and when said hydrogen gas
concentration is greater than a second threshold
concentration, said method further comprises:
maintaining a constant oxidant
stoichiometry when said hydrogen
35 concentration is decreasing; and

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measuring fuel cell voltage and
comparing said fuel cell voltage to a voltage
threshold value, and

5 if said fuel cell voltage exceeds
said voltage threshold value and said
hydrogen gas concentration is
increasing, decreasing the pressure of
said fuel stream;

10 if said fuel cell voltage is less
than said voltage threshold value, said
hydrogen gas concentration is
increasing, and oxidant mass flow rate
is less than a desired maximum, then
increasing said oxidant stoichiometry;
15 and

if said fuel cell voltage is less
than said voltage threshold value, said
hydrogen gas concentration is
increasing, and oxidant mass flow rate
20 is greater than or equal to a desired
maximum, then decreasing the pressure of
said fuel stream.

40. The method of claim 39 further
25 comprising regulating fluid pressure of said
oxidant and fuel streams to increase a pressure
differential between said oxidant and fuel
streams.

30 41. The method of claim 39 further
comprising regulating fluid pressure of said
oxidant and fuel streams to reduce a pressure
differential between said oxidant and fuel
streams.

35

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42. The method of claim 39 wherein said voltage threshold value is about 100 millivolts.

43. The method of claim 39 further
5 comprising ceasing operation of said fuel cell when electrical power output is less than electrical power demand.

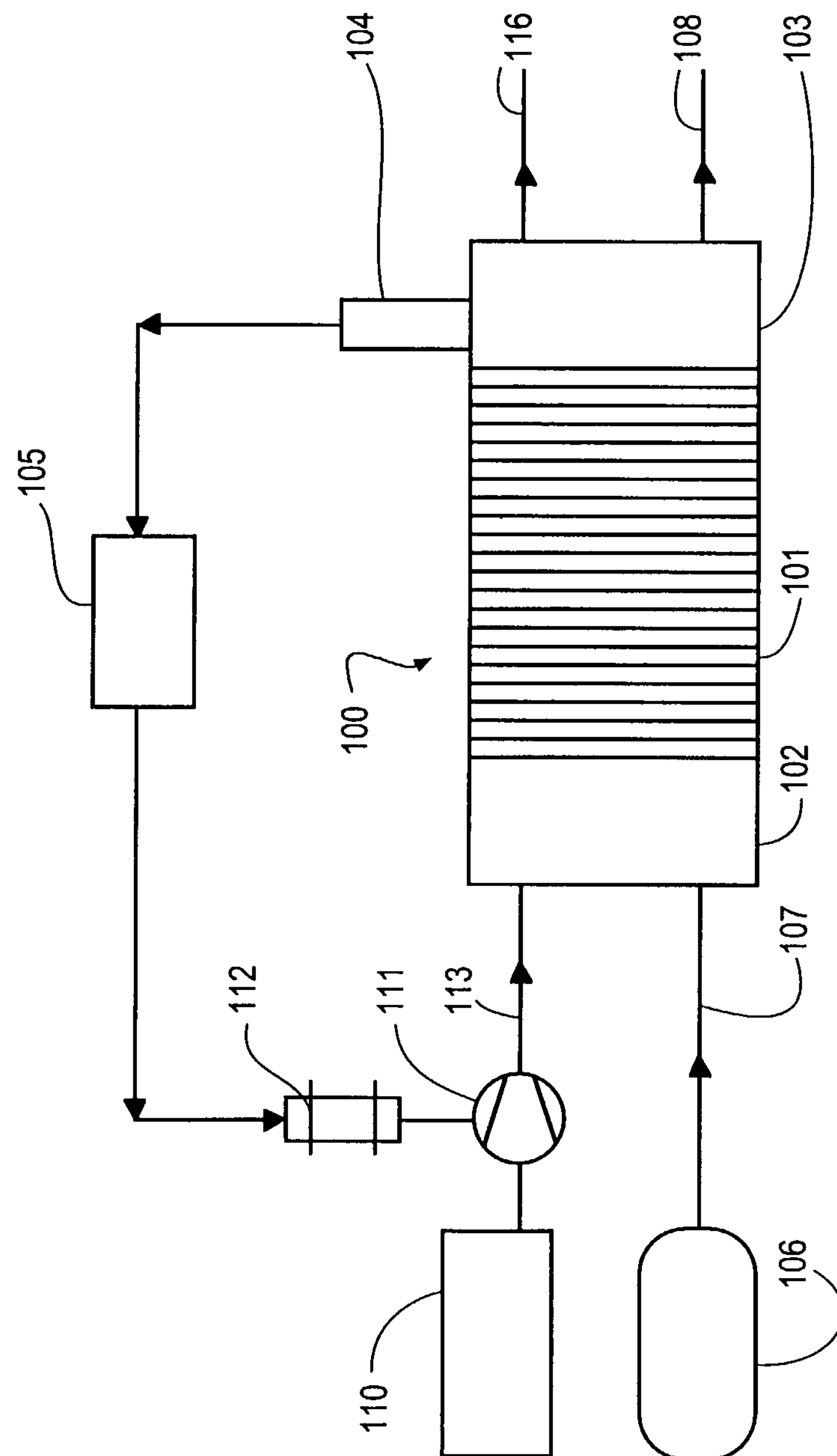
44. The method of claim 18 wherein said
10 monitoring is performed periodically.

45. The method of claim 26 wherein said oxidant stream mass flow rate is reduced until said second predetermined threshold concentration
15 of said hydrogen gas is detected and then increasing said oxidant mass flow rate by a predetermined percentage so that said oxidant stoichiometry is a predetermined percentage higher than about one.

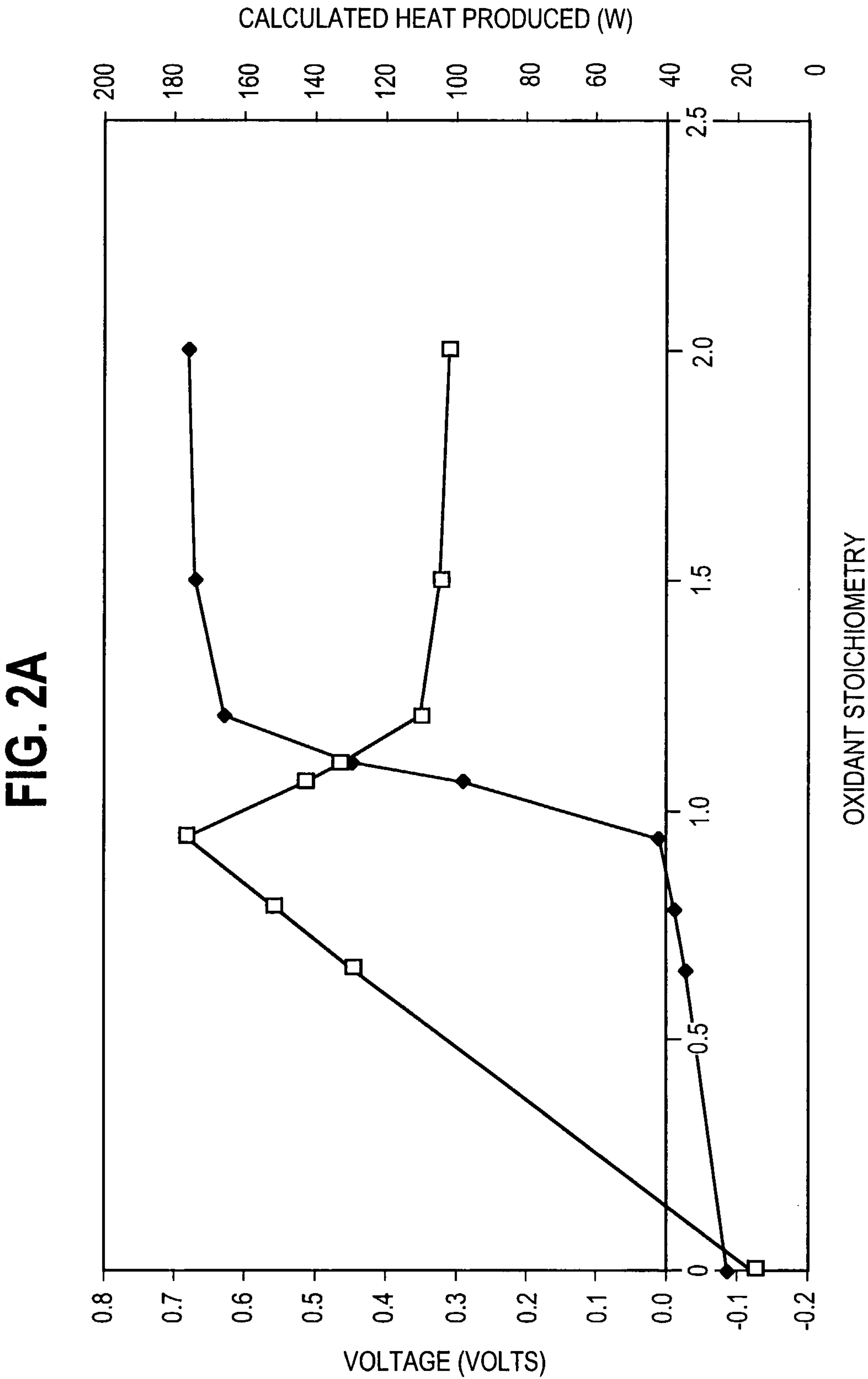
20

46. The method of claim 45 wherein said predetermined percentage is less than 50%.

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FIG. 1

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FIG. 2B

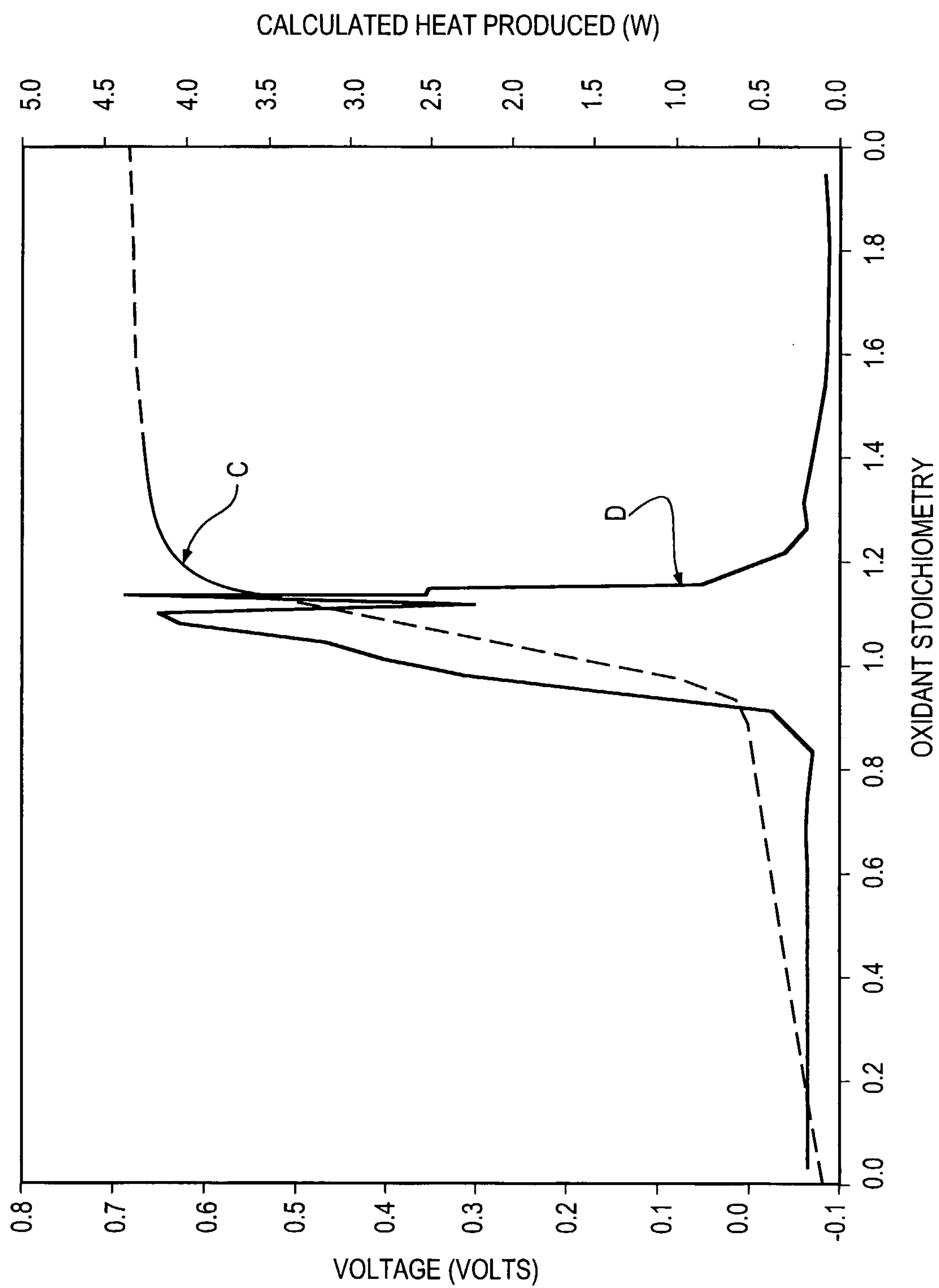
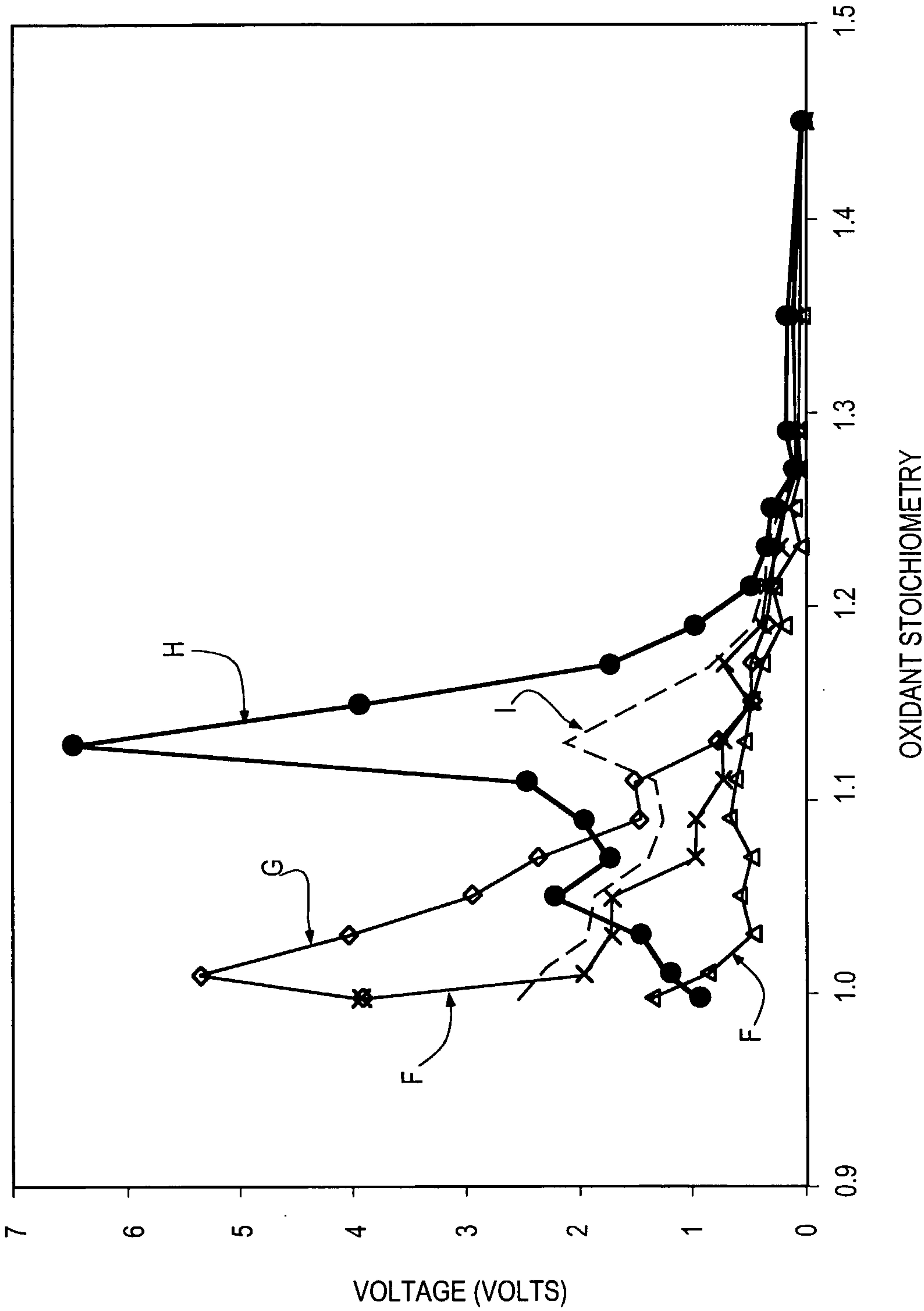


FIG. 2C



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FIG. 3A

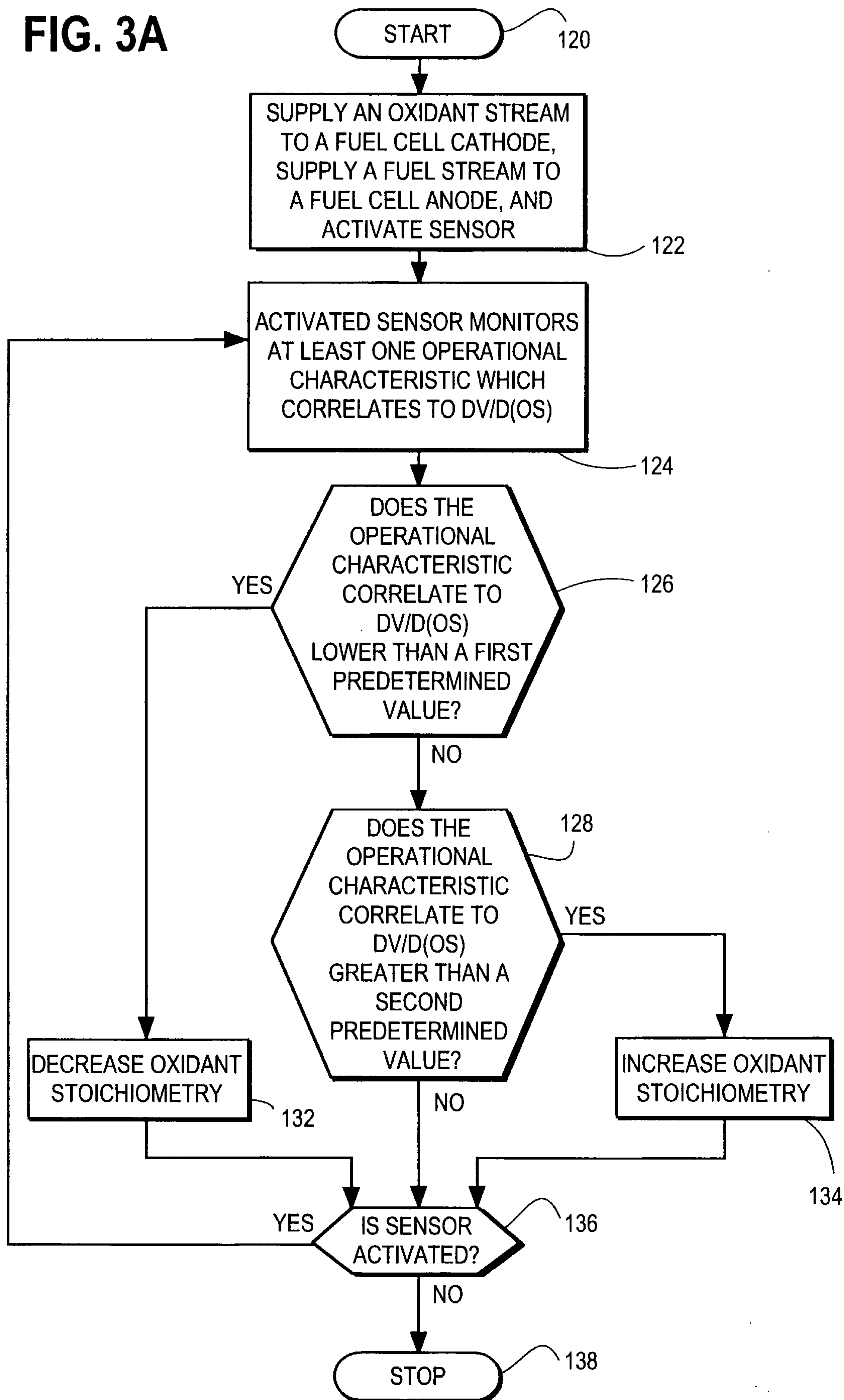
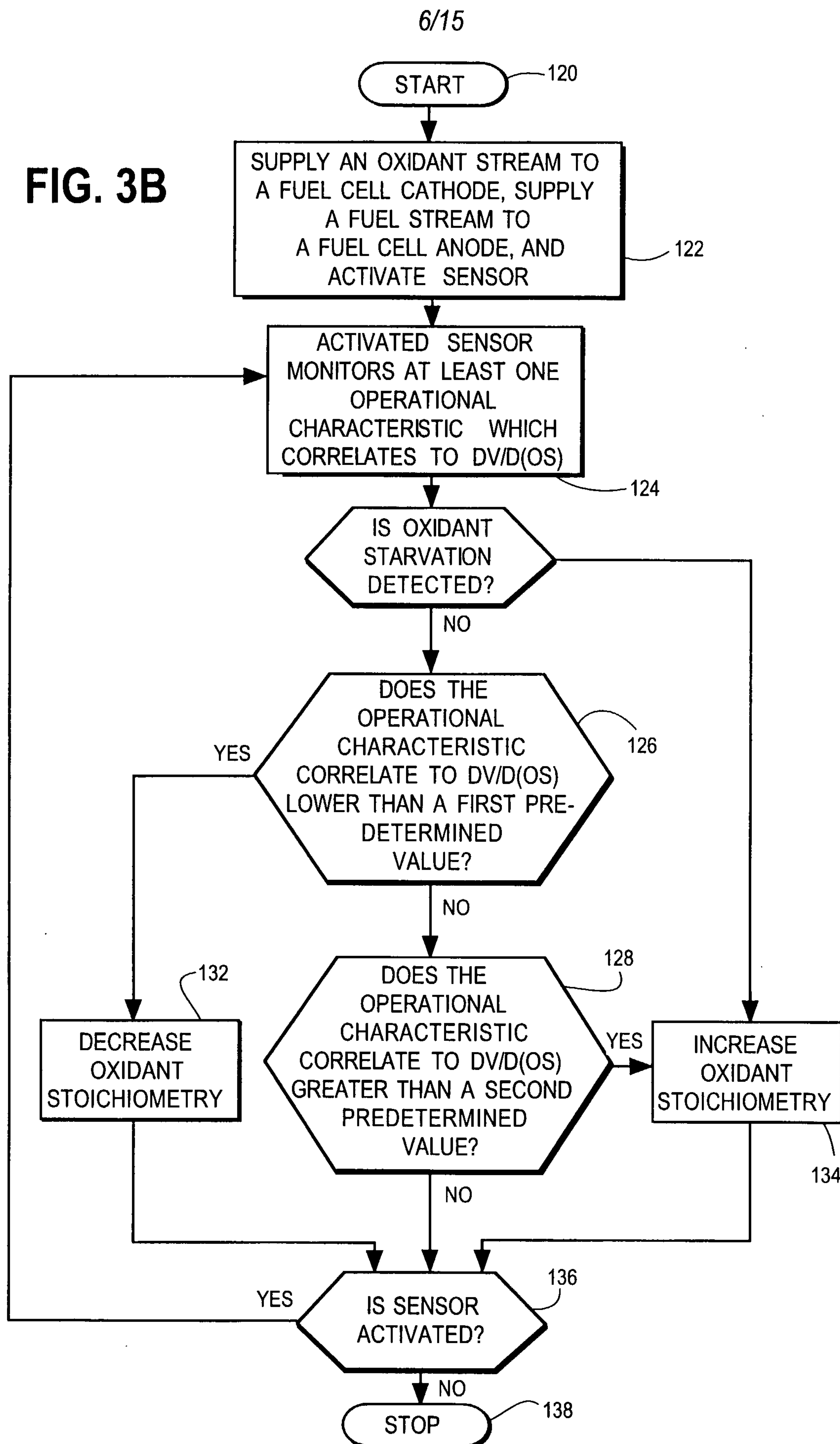
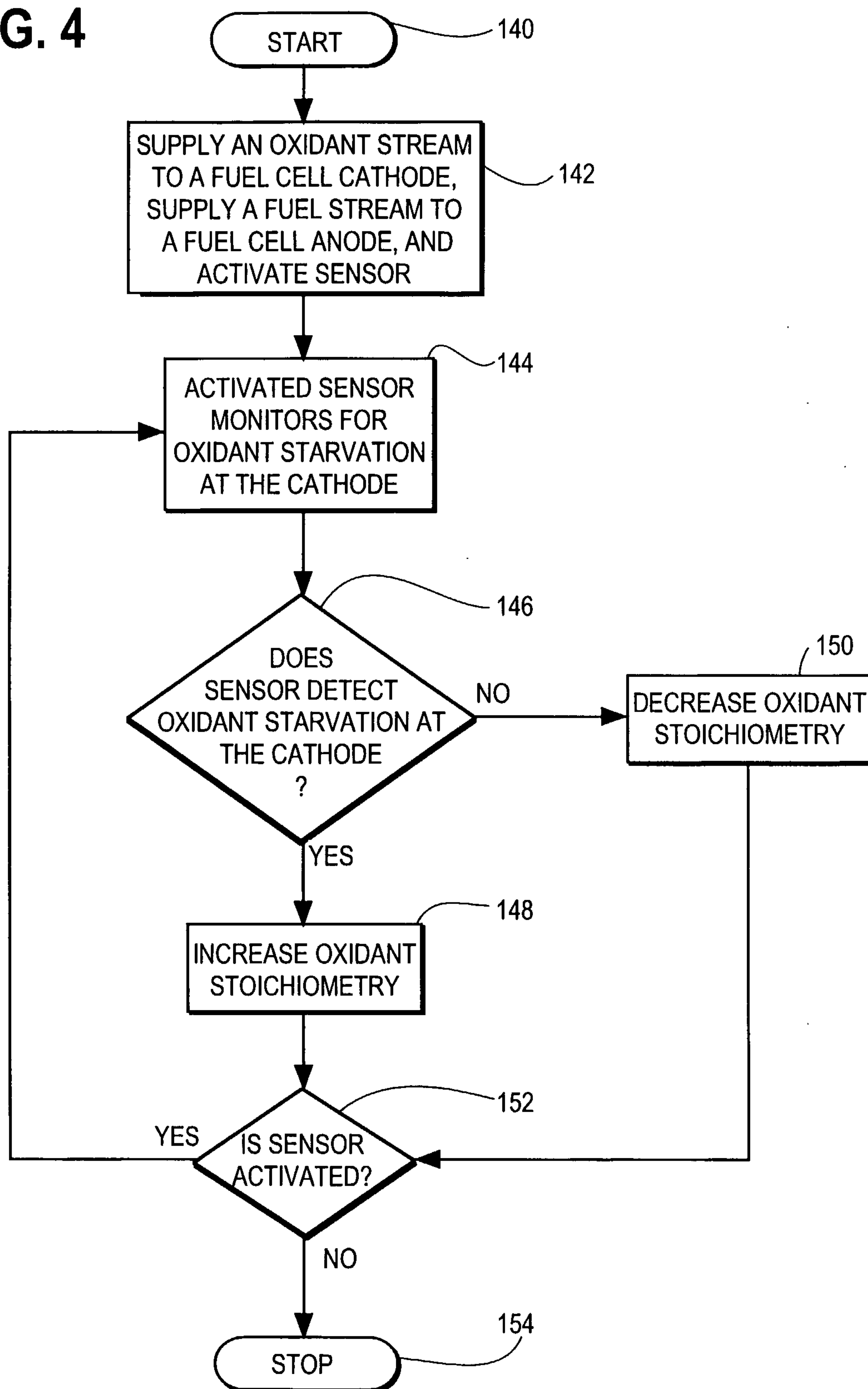


FIG. 3B

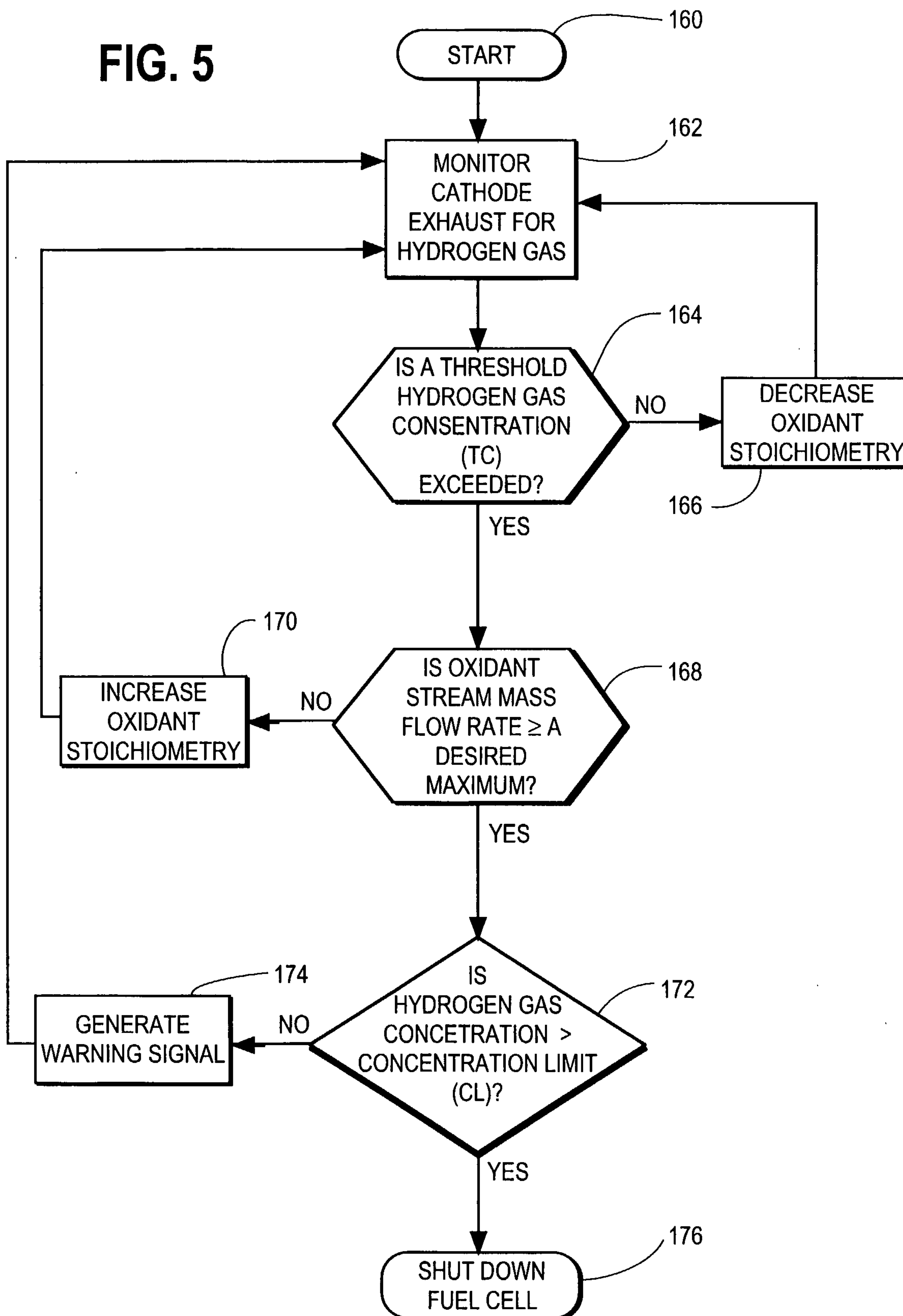


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FIG. 4

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FIG. 5



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FIG. 6

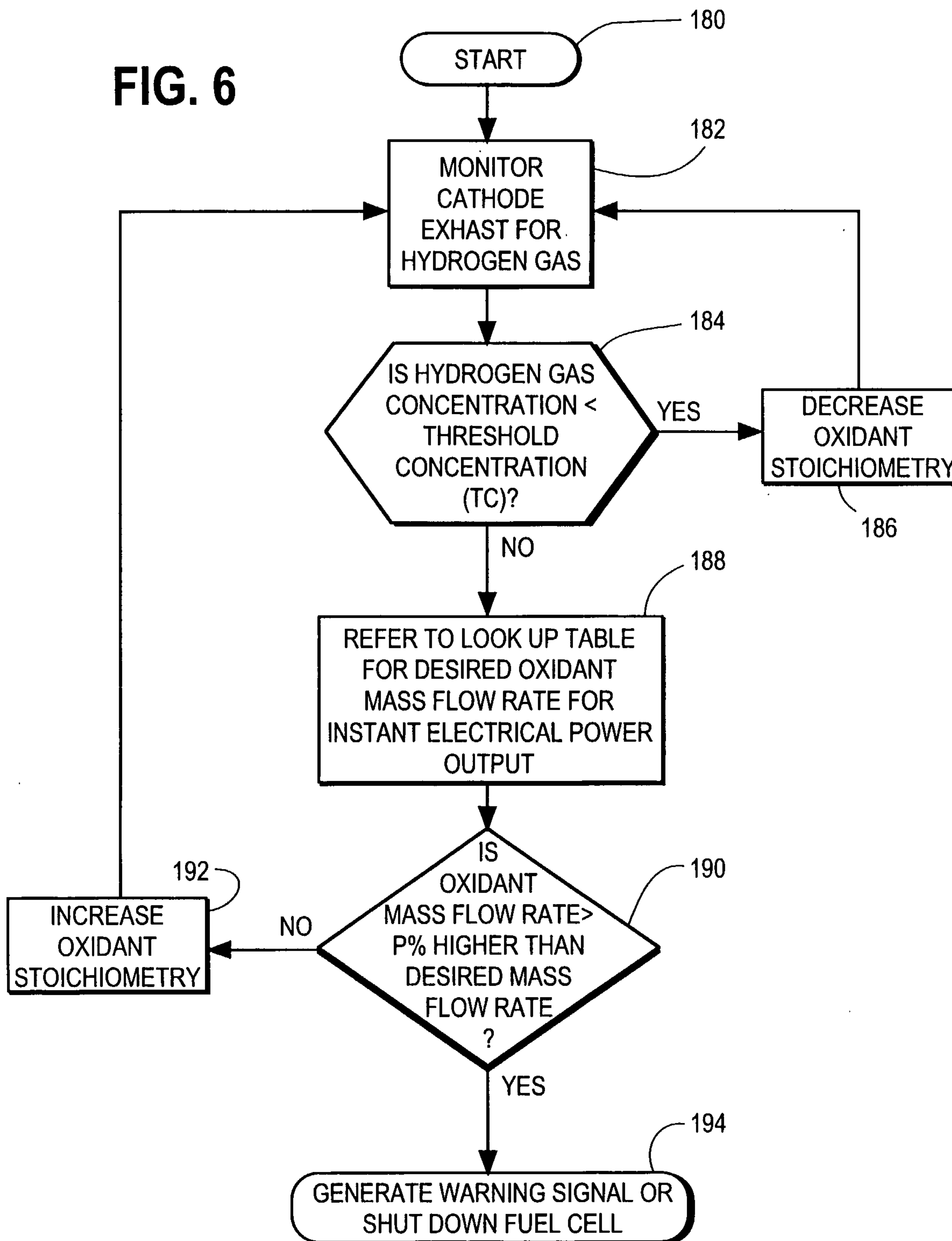
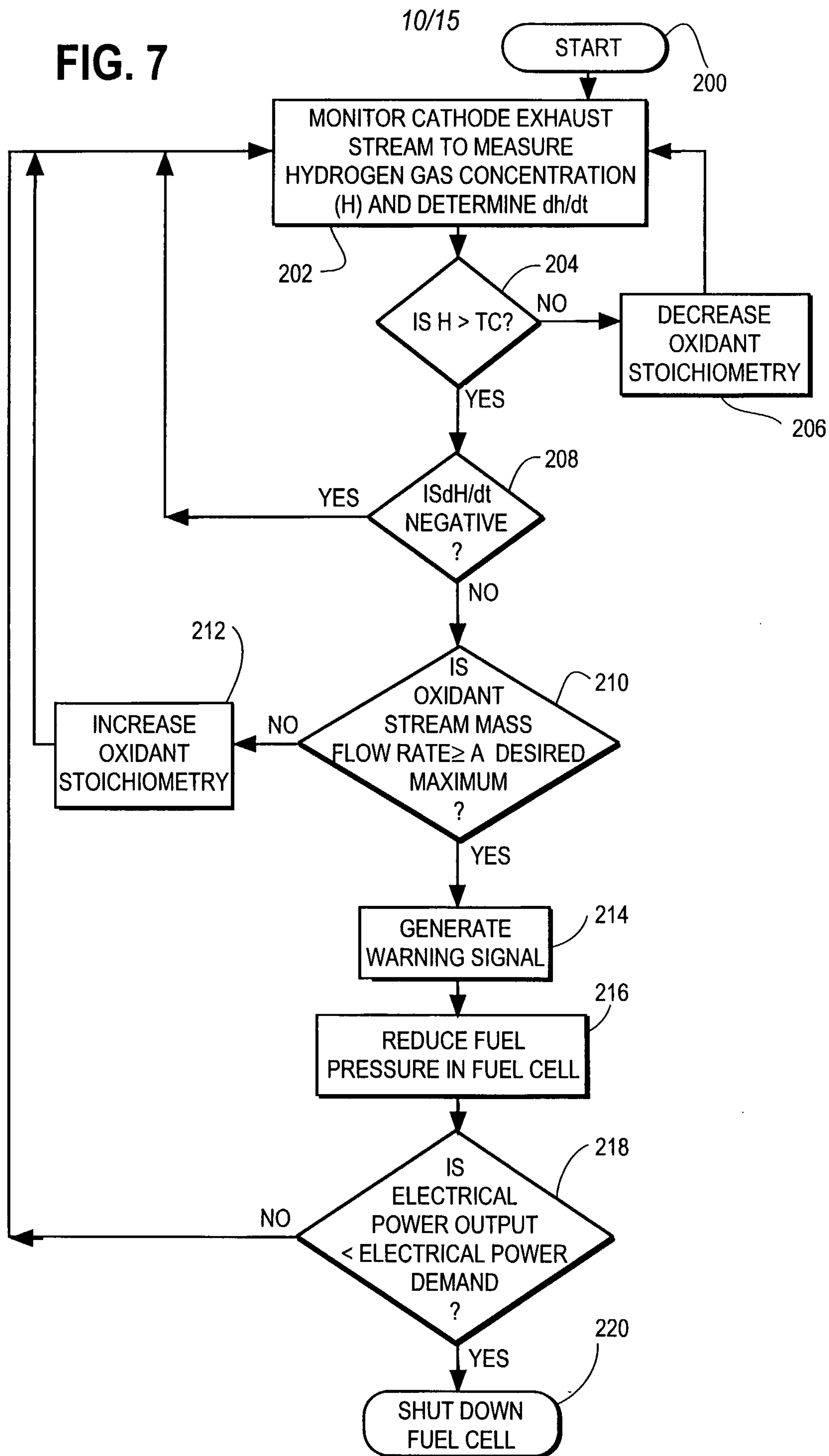


FIG. 7



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FIG. 8

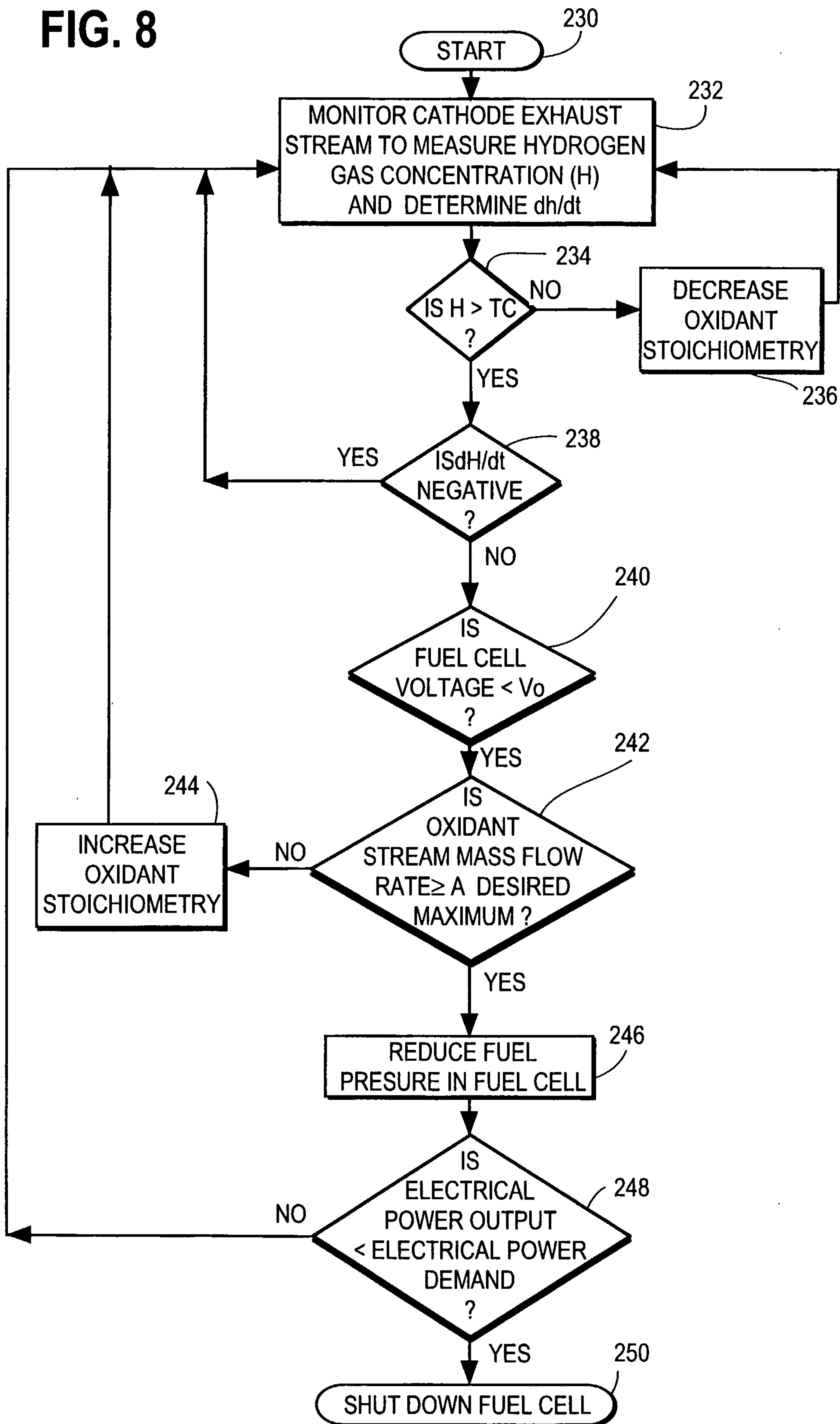
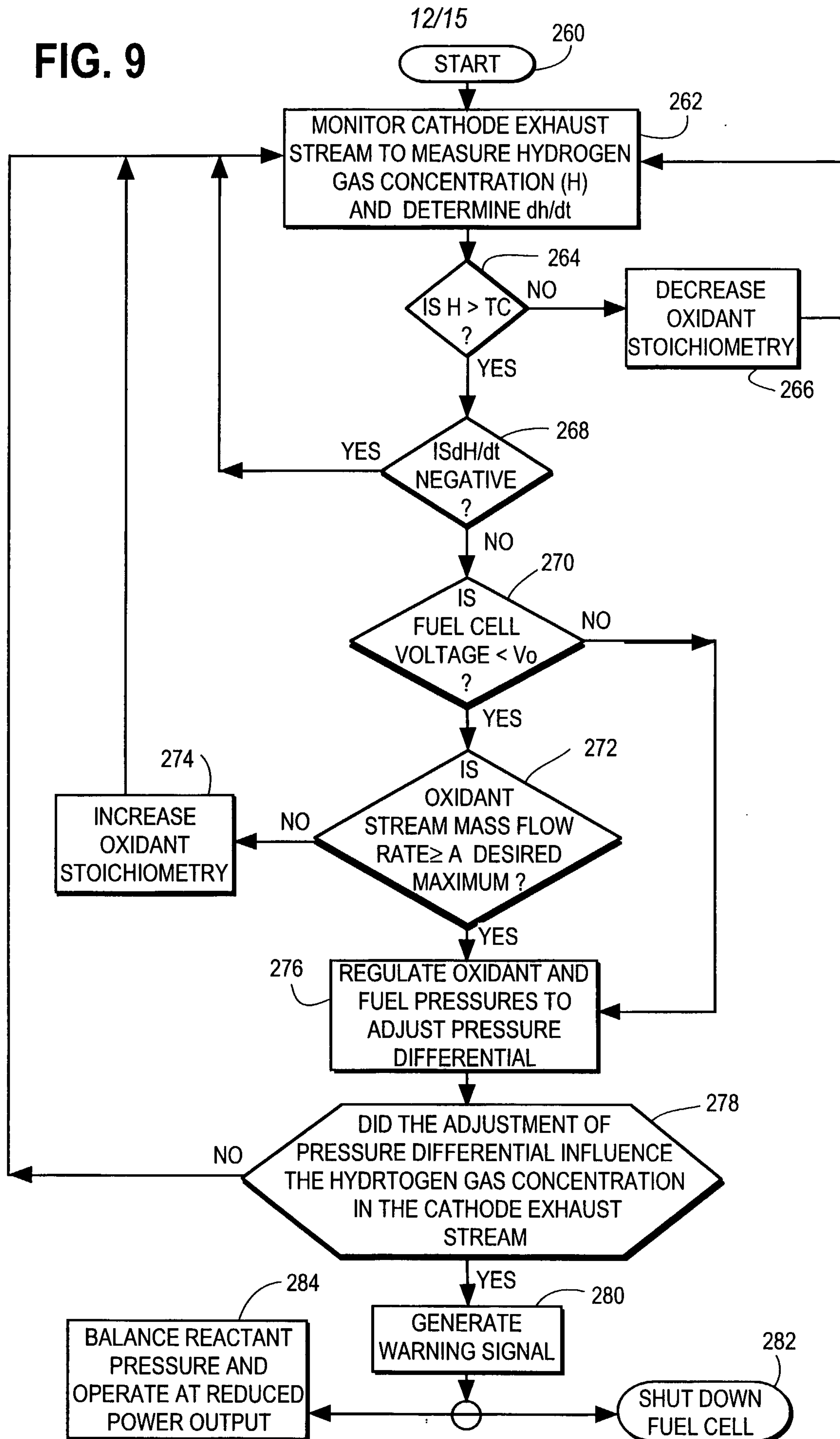
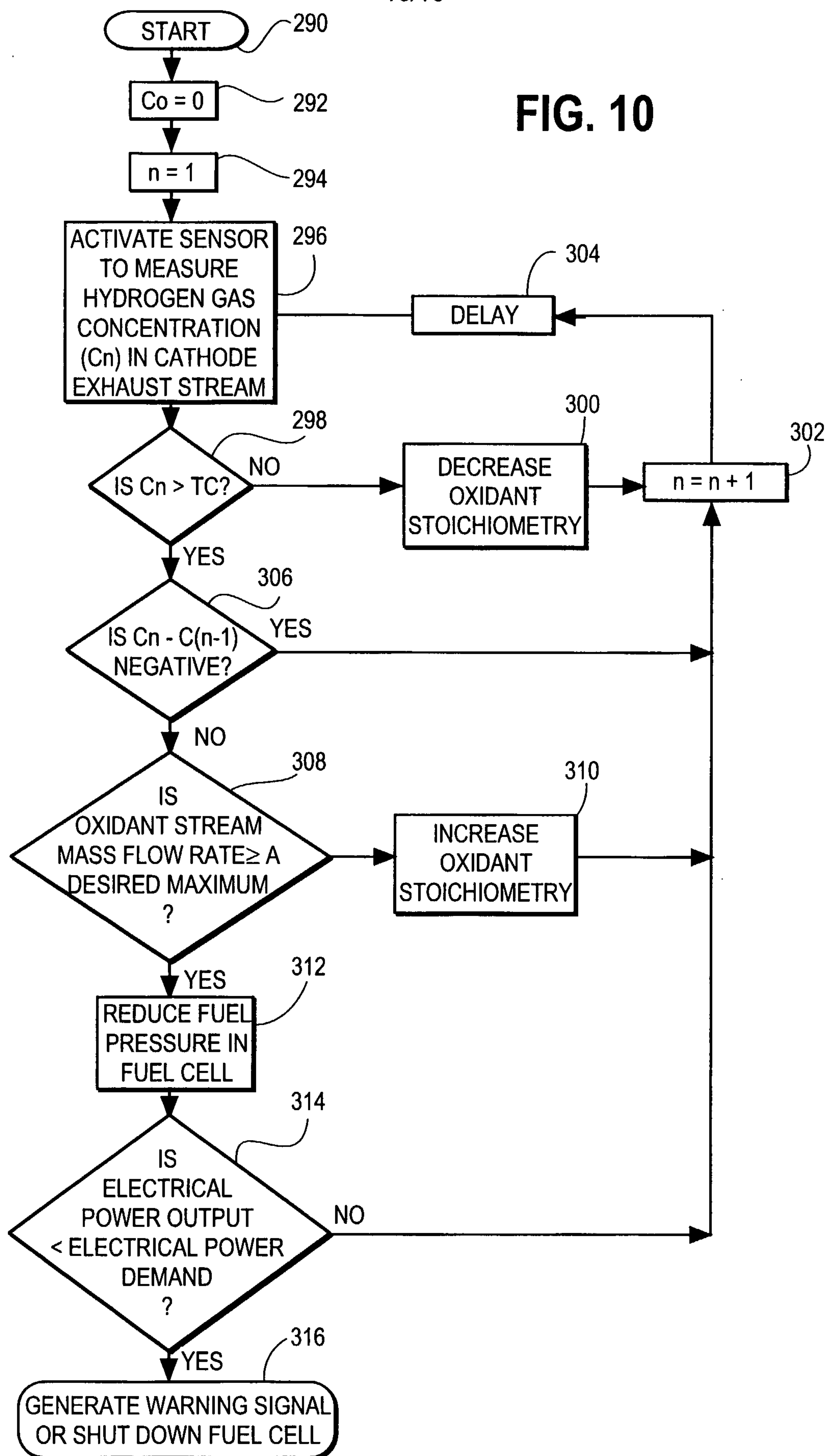


FIG. 9



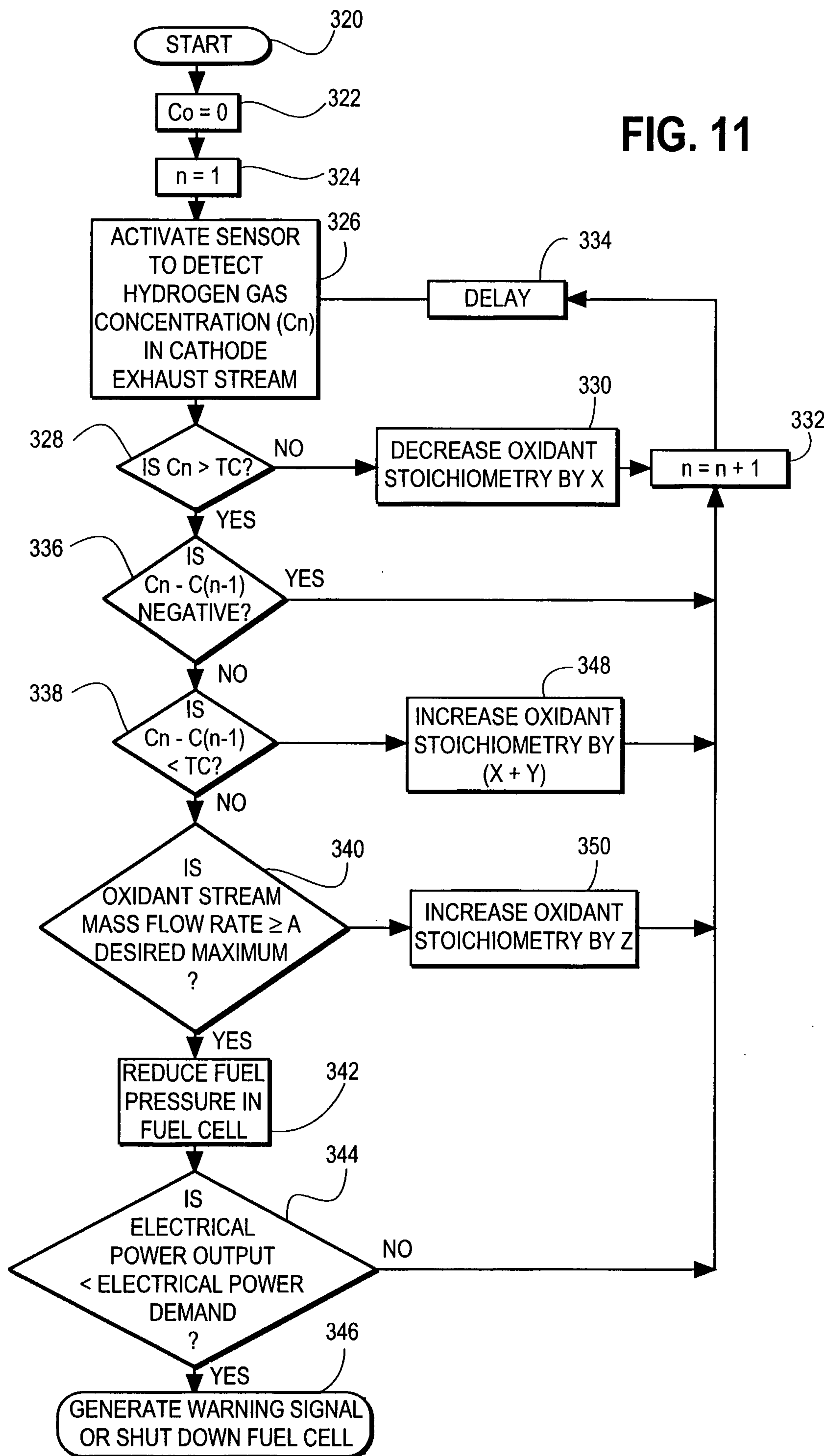
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FIG. 10



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FIG. 11



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FIG. 12

