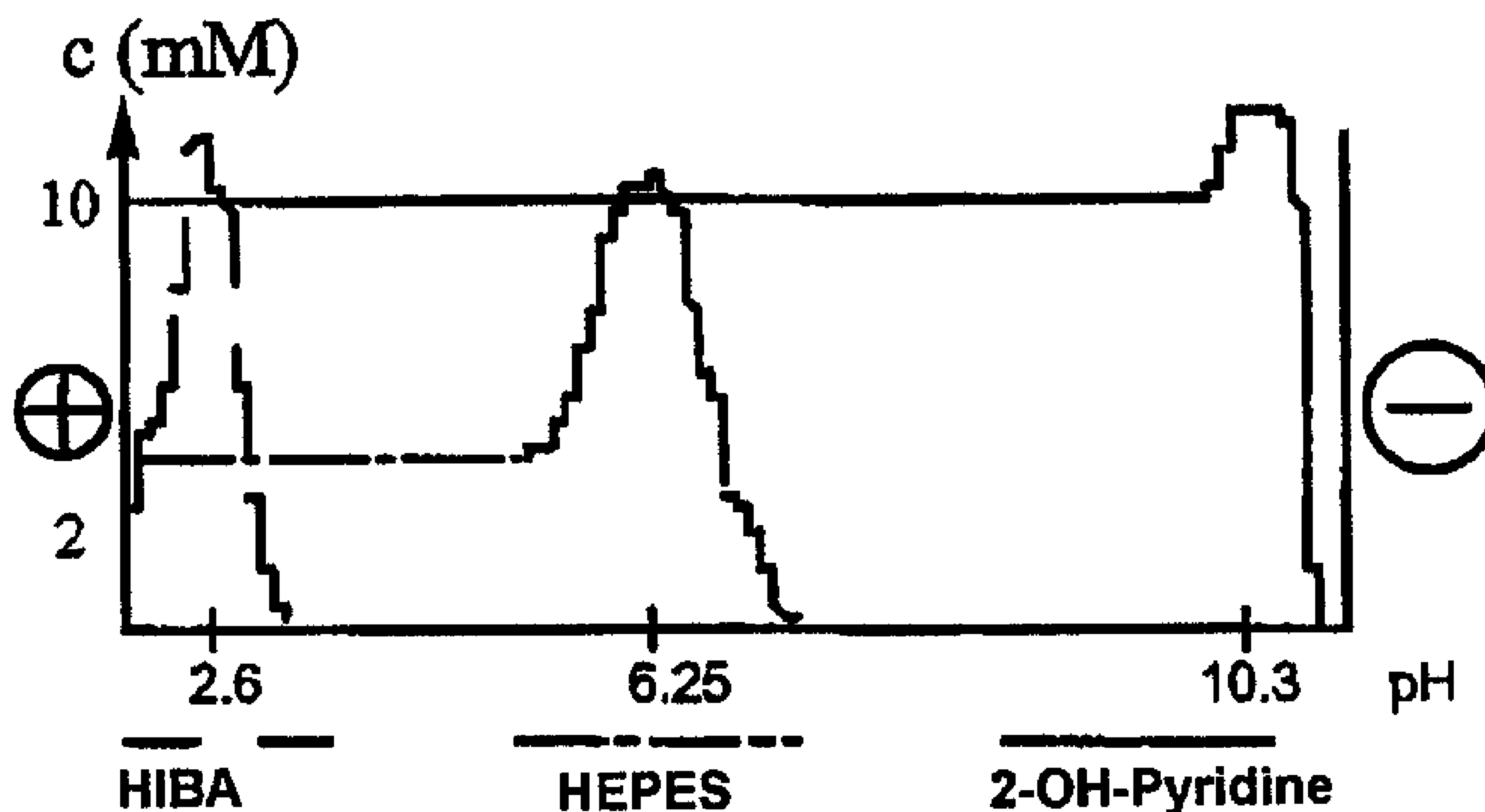




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(54) Titre : SUBSTANCE POUR ELECTROPHORESE ANALYTIQUE ET EXPERIMENTALE
(54) Title: MEDIUM FOR ANALYTIC AND PREPARATIVE ELECTROPHORESIS



(57) Abrégé/Abstract:

The invention relates to a medium for analytic and preparative electrophoresis comprising a content of acids and bases of different pKa values. Said medium contains at least two acids, whose pKa values (ΔpK_a) of adjacent acids are no greater than 1.0 but preferably range from 0.8 to 0.5, and contains at least two bases, whose pKa values range from approximately 1.5 to 11, whereby the difference from the pKa values (ΔpK_a) of adjacent bases is no greater than 1.0 but preferably ranges from 0.8 to 0.5.

ABSTRACT

The invention relates to a medium for analytic and preparative electrophoresis comprising a content of acids and bases of different pKa values. Said medium contains at least two acids, whose pKa values (ΔpK_a) of adjacent acids are no greater than 1.0 but preferably range from 0.8 to 0.5, and contains at least two bases, whose pKa values range from approximately 1.5 to 11, whereby the difference from the pKa values (ΔpK_a) of adjacent bases is no greater than 1.0 but preferably ranges from 0.8 to 0.5.

MEDIUM FOR ANALYTICAL AND PREPARATIVE ELECTROPHORESIS**FIELD OF THE INVENTION**

This present invention concerns a medium for analytical and preparative electrophoresis.

DESCRIPTION OF RELATED ART

Such aqueous solutions, which form a pH gradient upon high-voltage electrophoresis, are used on a large scale to analyze proteins and peptides. Such prior art solutions contain polyelectrolytes with amphoteric properties and are usually manufactured synthetically by means of condensation and/or polymerization of polyamines with unsaturated carboxylic acids. As a result of such synthesis, many different compounds with different molecular weights and a varying number of acid groups (carboxylic groups) and basic groups (amino and/or imino groups) are produced. Under the conditions prevalent during high-voltage electrophoresis, these polyelectrolytes arrange themselves, due to the different ratio between acidic and basic groups, in such a way that polyelectrolytes with particularly acidic properties concentrate in the proximity of the anode, and all other polyelectrolytes are arranged, in accordance with the "ranking" of acidic and basic properties, between the electrodes and also concentrated in the process. By the selective accumulation of polyelectrolytes in accordance with the acidic and/or basic properties of the individual species, a pH gradient is produced throughout the entire aqueous solution.

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By means of this pH gradient, amphoteric compounds such as proteins and peptides can be separated in accordance with the different acidic and/or basic properties. This analytical separation procedure is referred to in the literature as "isoelectric focusing". In case this procedure is carried out in a special gel (polyacrylamide), this procedure is referred to as PAGIEF

(polyacrylamide isoelectric focusing). PAGIEF is considered the standard procedure.

The unsaturated acids that are used in the synthesis of commercially available ampholytes, e.g. substituted acrylic acids, are highly toxic. Their residual content is considered to represent a significant hazard.

For that reason, the use of commercially available ampholytes for the cleaning of proteins for subsequent use in human applications is prohibited.

During ampholyte synthesis, a large number of chemical species of unknown structure is formed whose relative concentration may vary significantly across batches. This has an adverse impact on the reproducibility of analytical results.

In certain applications, "artificial bands" are reported as a result of undesired complex formation between the ampholyte and the analyte.

Partial interference with the analytical color reaction of proteins caused by ampholytes with a higher molecular weight has also been reported. An unusually high percentage of ampholytes with a molecular weight > 10 KD in products of certain manufacturers has been reported on several occasions.

In the case of preparative cleaning of proteins, the subsequent separation of the ampholyte is difficult and sometimes incomplete; the suspected cause is complex formation.

Commercial ampholytes have been found to be unsuitable for use as separating media for bioparticles, considering that bioparticles aggregate in these media.

SUMMARY OF THE INVENTION

The object of this invention is to provide aqueous solutions for isoelectric focusing (IF) which contain acids and bases with different pKa values which are not subject to the above-mentioned disadvantages and whose range of applications is therefore significantly wider.

These objects are achieved by an aqueous solution medium for analytical and preparative electrophoresis containing acids and bases with different pKa values, characterized in that the medium contains at least two acids whose pKa values are between approx. 3.0 and 12. Further, the difference of the pKa values (ΔpK_a) of neighboring acids is not larger than 1, and preferably between 0.8 and 0.5, and the medium contains at least two bases whose pKa values are between approx. 1.5 and 11, where the difference between the pKa values (ΔpK_a) of neighboring bases is not larger than 1, and preferably between 0.8 and 0.5.

The aqueous solution in accordance with this invention contains a mixture of acids and bases with graded "acidity" and/or "basicity" levels. The pKa values of the individual acids vary between pKa 3.5 to pKa 12, and the differences of the pKa values (ΔpK_a) of the "neighboring" acids should not exceed $\Delta pK_a = 1$. A value of ΔpK_a of less than 0.8 is considered ideal.

The pKa values of the bases used vary between pKa = 1.5 and pKa = 11. The differences between the values of the individual bases are ideal in case $\Delta pK_a < 0.8$ is fulfilled.

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The number of acids and/or bases used and their relative concentrations determine the range and the profile of the pH gradient. The minimum number is 2 acids and/or 2 bases each, and 3 acids and/or 3 bases are preferred. In the case of the wide pH range (pH 2.5 to 10), 11 acids and 13 bases with graded pKa values are used.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 provides a graphical illustration of the concentrations of typical acids and bases;

FIG. 2 graphically illustrates the relative concentrations of typical acids after electrophoretic migration; and

FIG. 3 graphically illustrates the relative concentrations of typical bases after electrophoretic migration.

DETAILED DESCRIPTION OF THE INVENTION

Figures 1 through 3, attached hereto, use a model of three typical acids and bases each to depict the electrophoretic migration and the resulting establishment of the pH gradient. Different concentrations of the acids hydroxyl butyric acid (HIBA: pKa = 3.8), HEPES (pKa = 7.5), and 2-hydroxy pyridine (pKa = 11.6), and the bases taurine

($pK_a = 1.5$), melamine ($pK_a = 5$), and ammediol ($pK_a = 9$) are contained in the original solution (see Figure 1). As a general rule, the relative concentrations of both very weak acids and very weak bases are significantly higher compared with the concentrations of the relatively strong acids and bases (approx. by a factor of five). The concentrations of acids and basis of "moderate" strength are increased by a factor of approx. 1.5 to a maximum of 3.

Figures 2 and 3 show the relative concentrations of the above-described acids and bases after electrophoretic migration. In the "state of equilibrium" shown herein, "pairs" of acids and bases are formed which may only ionize each other to a very limited degree (strong acid: HIBA and very weak base: taurine, weak acid: HEPES and weak base: melamine, and strong base: ammediol and very weak acid: 2OH pyridine), and as a result of thereof, the electrophoretic migration of the media's own substances comes to a stop. In the area of these pairs, a pH value is produced which corresponds to the arithmetic mean of the pK_a values of the acid and of the base. The pH values of these pairs are therefore as follows: HIBA/taurine solution, $pH = 2.6$; HEPES/melamine, $pH = 6.25$; 2-OH pyridine/ammediol, $pH = 10.3$. The pH range over which a pair is able to ensure a stable pH value is approx. one pH unit. All other acids and bases with different pK_a values form pairs which also ensure a stable pH value within one pH unit as long as the absolute concentrations of the individual acids and bases guarantee sufficient buffer capacity in the state of equilibrium.

The solutions in accordance with this invention consist of well-defined chemicals whose properties are

known in detail. The reproducibility of the entire mixture is ensured by the precision and/or accuracy of the scale used in its production. The substances used are non-toxic. Each of the substances used is "monofunctional", i.e. effects such as complex formation and/or biopolymer aggregation do either not occur at all or are at least significantly reduced compared with polyelectrolytes comprising several functional groups. The molecular weight of all individual species of the mixture is < 300 Dalton.

The aqueous solutions in accordance with this invention have already been successfully tested as separating media for biopolymers and bioparticles and have been found to be biocompatible. Considering that most individual species can be substituted by chemicals with similar pKa values, the special requirements of bioparticles in terms of separating media can be taken into account by changing the composition of the mixture. For the above reasons, official approval of such solutions for use in the cleaning of bioparticles and biopolymers (peptides and proteins) for subsequent application in humans appears possible.

Below, several examples of aqueous solutions in accordance with this invention with different pH intervals are presented. Please note that, during the manufacture of a medium for a narrower pH interval, the absolute concentrations of acids and bases are increased; usually, the differences in the pKa values are also reduced by using new constituents.

EXAMPLE 1
(pH Interval 3-10)

Acid	pKa	c (mMol)	Base	pKa	c (mMol)
Hydroxy					
Isobutyric Acid	4.0	1.5	Betaine	1.8	10
			2-Amino		
Pivalic Acid	5.0	5.0	Butyric Acid	2.3	10
Picolinic Acid	5.5	2	Creatine	2.7	10
			Nicotinic		
MES	6.1	4	Acid Amide	3.1	8
			Pip-4-		
			Carboxylic		
MOPSO	6.9	2	Acid	3.5	3
HEPES	7.5	2	EACA	4.4	2
EPPS	8.0	2	Melamine	5.0	2
			Hydroxyethyl		
TAPS	8.3	3	Pyridine	5.3	2
			Histidine	6.0	3
AMPSO	9.0	6	BISTRIS	6.5	3
			Hydroxyethyl		
CAPSO	9.6	8	Morpholine	7.1	3
			Triethanol		
CAPS	10.3	9	Amine	7.8	2
CABS	10.8	10	TRIS	8.3	2
2-Hydroxy-					
Pyridine	11.8	10	Ammediol	9.0	2

EXAMPLE 2
(pH Interval 4-6)

Acid	pKa	c (mMol)	Base	pKa	c (mMol)
Pivalic Acid	5.0	5	Creatine	2.7	30
			Glycl		
Picolinic Acid	5.5	10	Glycine	3.1	20
			Pip.-4-		
			Carboxylic		
MES	6.1	20	Acid	3.5	10
MOPSO	6.9	30	GABA	4.2	10
			EACA	4.4	5

EXAMPLE 3
(pH Interval 5.5 - 7.5)

Acid	pKa	c (mMol)	Base	pKa	c (mMol)
MOPSO	6.9	10	EACA	4.4	15
MOPS	7.2	10	Creatinine	4.8	15
			Hydroxyethyl		
HEPES	7.5	15	Pyridine	5.4	10
EPPS	8.0	15	β -Picoline	5.8	10
TAPS	8.3	20	BISTRIS	6.5	5

EXAMPLE 4(pH Interval 8-11.5)

Acid	pKa	C (mMol)	Base	pKa	c (mMol)
TAPS	8.3	3	Imidazol	7.0	15
			Triethanol		
AMPSO	9.0	5	Amine	7.8	12.5
CAPSO	9.6	7.5	TRIS	8.3	12.5
GABA	10.2	10	Ammediol	8.8	10
			Hydroxy		
			Ethyl		
			Ethylene		
EACA	10.6	10	Diamine	9.8	7.5
			Diethyl		
			Amino		
4-Hydroxy-			Ethylene		
Pyridine	11.2	7.5	Diamine	10.2	n 5
2-Hydroxy-			Triethyl		
Pyridine	11.6	7.5	Amine	10.8	5
			Dimethyl		
			Piperidine	11.2	3

EXAMPLE 5(pH Interval 2.5-5)

Acid	pKa	c (mMol)	Base	pKa	c (mMol)
Citric Acid	3.2	5	Betaine	1.8	20
			2-Amino		
HIBA	4.0	10	Butyric Acid	2.2	20
Isobutyric Acid	4.5	15	Creatine	2.7	15
			Glycyl		
Picolinic Acid	5.5	20	Glycine	3.1	10
			Pip.-4		
			Carboxylic		
MES	6.1	20	Acid	3.6	10
			GABA	4.2	5

In the specified media, mostly linear pH gradients are produced. The pH profile can be flattened in certain pH segments by modifying the relative concentrations of acids and bases. In the manufacture of the solutions in accordance with this invention, the order in which the individual compounds are added is usually irrelevant.

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What is claimed is:

1. A medium for electrophoresis, comprising at least two acids whose pKa values are between approximately 3.0 and approximately 12, and at least two bases whose pKa values are between approximately 1.5 and approximately 11, wherein for the majority of acids, an acid with a higher Pka value will have a higher relative and absolute concentration than an acid with a lower pKa value and wherein for the majority of bases, a base with a higher pKa value will have a lower relative and absolute concentration than a base with a lower pKa value.
2. The medium in accordance with claim 1, wherein for essentially all of the acids, an acid with a higher pKa value will have a higher relative and absolute concentration than an acid with a lower pKa value and wherein for the essentially all of the bases, a base with a higher pKa value will have a lower relative and absolute concentration than a base with a lower pKa value.
3. The medium in accordance with claim 1, wherein for all of the acids, an acid with a higher pKa value will have a higher relative and absolute concentration than an acid with a lower pKa value and wherein for all of the bases, a base with a higher pKa value will have a lower relative and absolute concentration than a base with a lower pKa value.
4. The medium in accordance with claim 1, wherein the medium contains at least three acids and at least three bases.

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5. The medium in accordance with claim 1, wherein the concentrations of the individual acids and bases are selected to provide sufficient buffer capacity in a state of equilibrium upon completion of the electrophoretic migration.

6. The medium in accordance with claim 1, wherein, in order to establish narrower pH intervals, the absolute concentrations of the acids and bases are increased to achieve the narrower pH intervals.

7. The medium in accordance with claim 1, wherein, in order to establish narrower pH intervals, the ΔpK_a values of the acids and bases are reduced through selection of the acids and bases to achieve the narrower pH intervals.

8. The medium in accordance with claim 1, wherein the molecular weights of both the acids and the bases used are less than 300 Dalton.

9. A medium for electrophoresis, comprising at least two acids whose pKa values are between approximately 3.0 and approximately 12, and at least two bases whose pKa values are between approximately 1.5 and approximately 11, wherein the relative concentrations of the weaker acids are 1.5 to 3 times higher than that of the strongest acid when the difference between the pKa values of the strongest acid and the weaker acid is in the range from 1 to 3, the relative concentrations of the weaker acids are 4 to 8 times higher than that of the strongest acid when the difference between the pKa values of the strongest acid and the weaker acid is in the range from 3 to 8, the relative concentrations of the weaker bases are 1.5

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to 3 times higher than that of the strongest base when the difference between the pKa values of the strongest base and the weaker base is in the range from 1 to 3, and the relative concentrations of the weaker bases are 4 to 8 times higher than that of the strongest base when the difference between the pKa values of the strongest base and the weaker base is in the range from 3 to 8.

10. The medium in accordance with claim 9, wherein the medium contains at least three acids and at least three bases.

11. The medium in accordance with claim 10, wherein ΔpK_a of neighboring acids is not larger than 1, and wherein ΔpK_a of neighboring bases is not larger than 1.

12. The medium in accordance with claim 11, wherein ΔpK_a of neighboring acids is between 0.8 and 0.5.

13. The medium in accordance with claim 11, wherein ΔpK_a of neighboring bases is between 0.8 and 0.5.

14. A medium for electrophoresis, comprising at least two acids whose pKa values are between approximately 3.0 and approximately 12, and at least two bases whose pKa values are between approximately 1.5 and approximately 11, wherein for the majority of acids, an acid with a higher pKa value will have a higher relative and absolute concentration than an acid with a lower pKa value and wherein for the majority of bases, a base with a higher pKa value will have a lower relative and absolute concentration than a base with a lower pKa value, wherein the relative concentrations of the weaker acids are

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1.5 to 3 times higher than that of the strongest acid when the difference between the pKa values of the strongest acid and the weaker acid is in the range from 1 to 3, the relative concentrations of the weaker acids are 4 to 8 times higher than that of the strongest acid when the difference between the pKa values of the strongest acid and the weaker acid is in the range from 3 to 8, the relative concentrations of the weaker bases are 1.5 to 3 times higher than that of the strongest base when the difference between the pKa values of the strongest base and the weaker base is in the range from 1 to 3, and the relative concentrations of the weaker bases are 4 to 8 times higher than that of the strongest base when the difference between the pKa values of the strongest base and the weaker base is in the range from 3 to 8.

15. The medium in accordance with claim 14, wherein the medium contains at least three acids and at least three bases.

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