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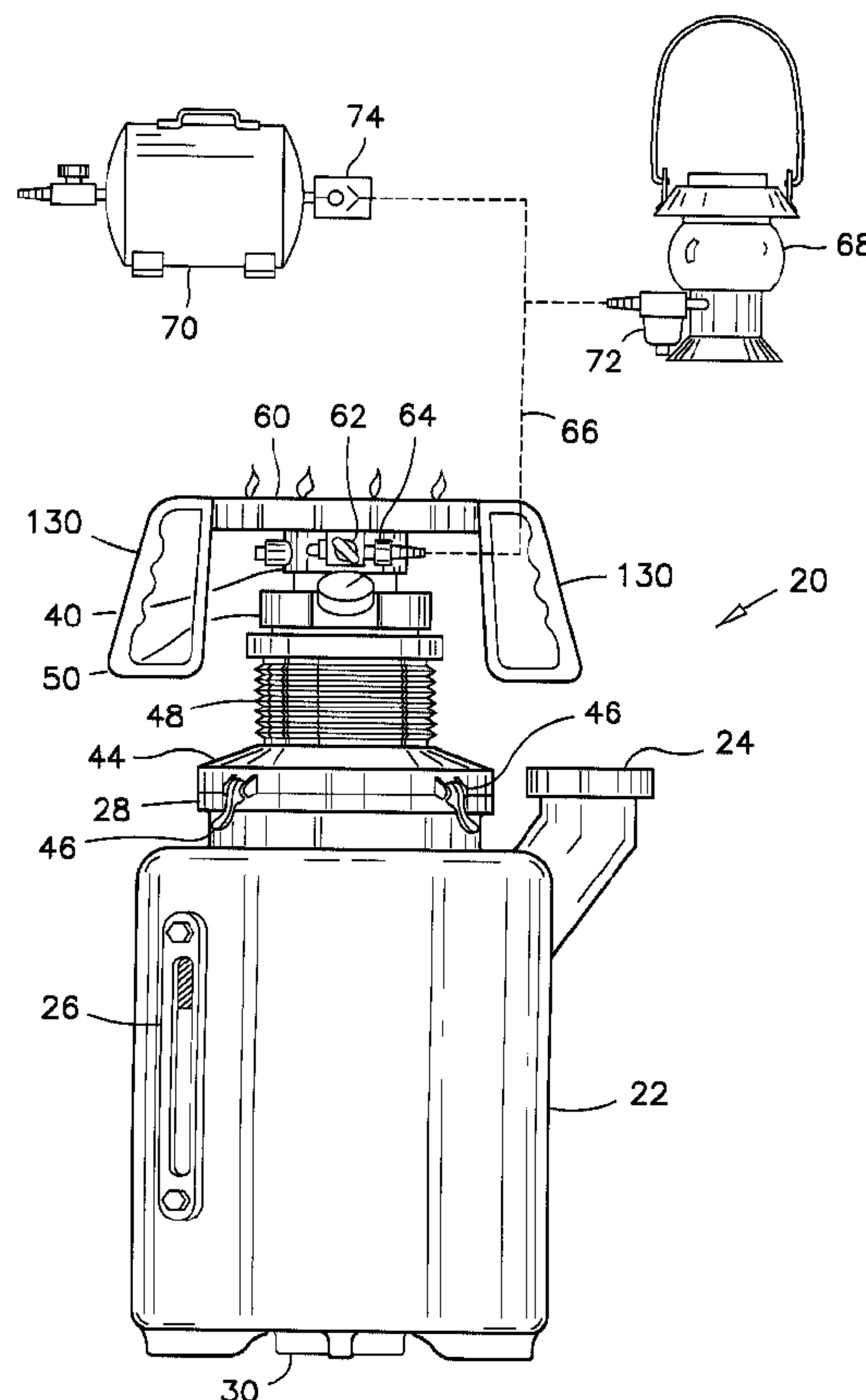
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(54) Titre : PRODUCTION D'HYDROGENE A PARTIR DE L'ALUMINIUM, EAU ET HYDROXIDE DE SODIUM EN TANT QUE CATALYSEUR

(54) Title: HYDROGEN PRODUCTION FROM ALUMINUM, WATER, AND SODIUM HYDROXIDE AS CATALYST



(57) Abrégé/Abstract:

The process for producing hydrogen gas according to the present invention consists of reacting aluminum with water in the presence of sodium hydroxide as a catalyst. An apparatus for carrying out the method is also described. The apparatus comprises an expandable container wherein the pressure and temperature of the reaction causes the container to expand and



(57) **Abrégé(suite)/Abstract(continued):**

contract to control the degree of immersion of a fuel cartridge in water and consequently to control the intensity and duration of the reaction.

**TITLE: HYDROGEN PRODUCTION FROM ALUMINUM,
WATER, AND SODIUM HYDROXIDE AS CATALYST**

ABSTRACT OF THE DISCLOSURE

The process for producing hydrogen gas according to the present invention consists of reacting aluminum with water in the presence of sodium hydroxide as a catalyst. An apparatus for carrying out the method is also described. The apparatus comprises an expandable container wherein the pressure and temperature of the reaction causes the container to expand and contract to control the degree of immersion of a fuel cartridge in water and consequently to control the intensity and duration of the reaction.

**TITLE: HYDROGEN PRODUCTION FROM ALUMINUM,
WATER, AND SODIUM HYDROXIDE AS CATALYST**

FIELD OF THE INVENTION

This invention relates to the production of hydrogen gas from aluminum, water, and sodium hydroxide as catalyst, and to an apparatus for carrying out the method.

BACKGROUND OF THE INVENTION

5 Generally speaking, it is known that under certain conditions, aluminum reacts with water to generate hydrogen and heat. It is also known, however, that this type of reaction is not sustainable at ambient temperature. It is believed that a protective oxide layer forms on a metal surface in contact with water at ambient temperature and hampers the
10 reaction. Therefore, it has been accepted by those skilled in the art that the use of aluminum in a reaction with water to generate heat and hydrogen gas requires that the protective oxide layer is efficiently and continuously removed, and that the reaction is kept at an elevated temperature.

 A number of hydrogen generators have been developed in the past.
15 The following patent documents constitute a good inventory of the devices and methods of the prior art in the field of hydrogen gas generation using the reaction of aluminum or alloys of aluminum with water.

US **909,536** issued on Jan. 12, 1909, and US **934,036** issued on Sept. 14, 1909, both issued to G. F. Brindley et al. These documents disclose
20 several compositions for generating hydrogen. The compositions comprise

any metal which can form an hydroxide when it is brought into contact with a solution of a suitable hydroxide. For example, aluminum is reacted with sodium hydroxide to release hydrogen and produce sodium aluminate.

5 **US 2,721,789**, issued on Oct. 25, 1955 to Q.C. Gill. This document discloses the structure of an hydrogen generator for reacting water with a measured dry charge of aluminum particles and flakes of sodium hydroxide. The reaction releases hydrogen gas and produces sodium aluminate.

10 **US 3,554,707** issued on Jan. 12, 1971 to W.A. Holmes et al. This document discloses a gas generator having bellows to raise or lower the level of water in response to the pressure inside the generator. As the level of water drops, the contact surface between the fuel cartridge and the water is lost and the reaction is terminated.

15 **US 3,957,483** issued on May 18, 1976 to M. Suzuki. This patent discloses a magnesium composition which produces hydrogen upon contact with water. The preferred magnesium composition comprises magnesium, and one or more metals selected from the group consisting of iron, zinc, chromium, aluminum and manganese.

20 **US 3,975,913** issued on Aug. 24, 1976 to D.C. Erickson. This document discloses a hydrogen generator wherein molten aluminum is reacted with water. The generator is kept at a very high temperature to keep the metal in a molten condition.

25 **US 4,643,166** issued on Feb. 17, 1987, and
 US 4,730,601 issued on Mar. 15, 1988 both to H.D. Hubele et al. These documents disclose the structure of a fuel cell for producing heat energy

and hydrogen gas. The device has a reaction chamber containing a fuel composition that is reactive with water. The fuel composition includes a main fuel part of magnesium and aluminum in a molar ratio of 1:2, and the second part is composed of lithium hydride, magnesium and aluminum in equal molar ratio.

US 4,670,018 issued on June 2, 1987, and

US 4,769,044 issued on Sept. 6, 1988, both to J.H. Cornwell. These documents describe a log made of compressed wood waste and paper. The log is coated with aluminum particles. Upon burning, the aluminum particles react with moisture in the log to emit heat due to the generation of hydrogen gas.

US 4,752,463 issued on June 21, 1988 to K. Nagira et al. This document discloses an alloy which reacts with water for producing hydrogen gas. The alloy material comprises essentially aluminum and 5 to 50% tin.

US 5,143,047 issued on Sept. 1, 1992 to W.W. Lee. This document discloses an apparatus and a method for generating steam and hydrogen gas. In this apparatus, an aluminum or aluminum alloy powder is reacted with water to generate hydrogen gas. An electric power source is used to start the reaction. The electric power source is used to explode an aluminum conductor and to disperse pieces of molten aluminum into a mixture of water and aluminum powder. A heat exchanger is provided to extract useful heat.

US 5,867,978 issued on Feb. 9, 1999 to M. Klanchar et al. This document discloses another hydrogen gas generator using a charge of fuel selected from the group consisting of lithium, alloys of lithium and aluminum. The charge of fuel is molten and mixed with water to generate hydrogen gas.

JP 401,208,301 issued to Mito on Aug. 22, 1989. This document discloses a process for producing hydrogen. Aluminum is reacted with water under an inactive gas or a vacuum to produce hydrogen gas.

CA 2,225,978 published on June 29, 1999 by J. H. Checketts. This patent application discloses a hydrogen generation system wherein a coating on reactive pellets is selectively removed to expose the reactive material to water for producing hydrogen gas on demand. In one embodiment, aluminum and sodium hydroxide are reacted with water to release hydrogen gas and produce sodium aluminate.

Various other processes to produce hydrogen gas have been described in the art, as reacting water with magnesium, sodium, potassium, lithium, calcium, iron, zinc or steel.

Although the hydrogen production processes of the prior art deserve undeniable merits, it is believed that the catalytic reaction of aluminum and water, using sodium hydroxide as the catalyst, to release hydrogen gas from water at room temperature has never been anticipated or observed and disclosed by prior inventors. It is also believed that the prior art is short of suggestion with regards to a hydrogen production process which can be improvised in a home workshop using common materials and equipment, to generate heat and light during a power outage for example.

Concerning hydrogen generators, the prior art discloses a number of hydrogen generators for use with fuel cells or as heat sources for thermal engines for examples. These generators are believed to be complicated and precarious to operate by untrained individuals. These hydrogen generators are believed to be designed for use by scientists and other professionals working under laboratory conditions.

As such , it will be appreciated that there continues to be a need for a production process and for an apparatus for generating hydrogen gas and heat using a simple reaction which can be started at room temperature and carried out safely by ordinary persons not having a formal education in chemistry and chemical processes.

SUMMARY OF THE INVENTION

Broadly stated, the process for producing hydrogen gas according to the present invention consists of reacting aluminum with water in the presence of sodium hydroxide as a catalyst. This process is advantageous for being carried out at room temperature and for producing large quantities of heat and hydrogen gas at high purity.

In accordance with another feature of the present invention, there is provided a process for producing heat, light and hydrogen gas. The process comprises the steps of providing an expandable receptacle; partly filling the expandable receptacle with water and introducing an aluminum element and a catalyst in the water. The process also comprises the steps of partly sealing the expandable receptacle and reacting the aluminum element with the water. Then, the expandable receptacle is expanded and contracted in response to more or less pressure therein, and by the same action, the fuel element is emerged out or immersed into the water. This method is advantageous for providing the ability to control the intensity of the reaction between the water and the aluminum element in response to the pressure generated inside the expandable receptacle by the reaction.

In another aspect of the present invention, the aluminum element comprises a coiled strip of aluminum having several layers set vertically in the water. As hydrogen gas is generated between the layers of the coiled strip of aluminum, the hydrogen gas raising to the top of the water causes a partial vacuum between the layers of the coiled strip, to absorb more water through the bottom of the coiled strip, thereby promoting an effective wetting of the aluminum element.

In yet another aspect of the present invention, there is provided an apparatus for producing heat, light and hydrogen gas. The apparatus comprises essentially an expandable receptacle having an upper end, a central portion and a fuel element suspended to the upper end and inside the central portion. The apparatus also has means for raising and lowering the fuel element in the central portion in response to more or less pressure inside the expandable receptacle, respectively.

The apparatus according to the present invention uses the pressure and temperature of a reaction occurring between a fuel element and the water contained therein to control the degree of immersion of a fuel element in the water and consequently to control the intensity and duration of the reaction between the fuel element and the water.

In yet a further feature of the present invention, the apparatus comprises a timer mechanism and latch means responsive to the timer mechanism for timely raising the fuel element out of the water contained in the receptacle.

The processes and apparatus according to the present invention are practical and safe for use by the general public to generate heat, light and

hydrogen gas in power outage situations for example, or in remote locations where electricity is not available. Furthermore, the method and apparatus according to the present invention use aluminum waste readily available in domestic garbage and metal working shops, to promote recycling and energy conservation.

In accordance with yet another aspect of the present invention, there is provided a process for producing alumina, comprising the step of reacting aluminum with water in the presence of a catalyst wherein the catalyst is sodium hydroxide. This process is advantageous for extracting available energy from a reaction between aluminum waste and water, and for simultaneously producing a basic material which can be reused for manufacturing new aluminum.

Although, the utility of the processes and apparatus of the present invention can be appreciated by the general public, it is also believed that the processes and apparatus of the present invention will find advantageous applications in other more scientific fields, such as the fields of fuel cells, internal combustion engines, thermal engines, heating systems and lighting appliances.

Other advantages and novel features of the present invention will become apparent from the following detailed description of the preferred embodiment.

BRIEF DESCRIPTION OF THE DRAWINGS

A preferred embodiment of the present invention selected by way of examples will now be described with reference to the accompanying drawings, in which:

FIG. 1 is a side view of the preferred energy production apparatus, also referred to herein as the hydrogen generator;

FIG. 2 is a cross-section view of the energy production apparatus illustrating a mode of operation thereof when the fuel cartridge is entirely immersed in water;

FIG. 3 is another cross-section view of the energy production apparatus with the fuel cartridge in a raised position when pressure inside the apparatus force the bellows of the apparatus to expand;

FIG. 4 illustrates yet another cross-section view of the energy production apparatus with the timer mechanism in an unlatched mode causing a spring to pull the cartridge out of the water;

FIG. 5 is a schematic diagram of the preferred gas handling manifold and a burner plate mounted on the energy production apparatus;

FIG. 6 is a side view of the upper fuel support portion of the energy production apparatus;

FIG. 7 illustrates a side view of a preferred burner plate and an optional heat storage device for use with the energy production apparatus;

FIG. 8 is a top view of the preferred timer mechanism for use with the energy production apparatus;

FIG. 9 is a partial cross-section view through the timer mechanism along line 9-9 in **FIG. 8.**;

FIG. 10 illustrates a first arrangement for a fuel cartridge for use with the energy production apparatus;

FIG. 11 illustrates a second arrangement for a fuel unit for use with the energy production apparatus;

5 **FIG. 12** illustrates one form for the fuel pellet for use with the energy production apparatus;

FIG. 13 illustrates a third arrangement for a fuel unit for use with the energy production apparatus.

10 **FIG. 14** illustrates a graph of temperature over time for a typical hydrogen gas production reaction.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENT

15 While this invention is susceptible of embodiments in many different forms, there is shown in the drawings and will be described in details herein a specific embodiment of the method and apparatus according to the present invention, with the understanding that the present disclosure is to be considered as an example of the principles of the invention and is not intended to limit the invention to the embodiment illustrated.

20 The production of hydrogen gas according to the present invention is obtained by a reaction of aluminum with water in the presence of sodium hydroxide (NaOH) as a catalyst. The reaction produces a large amount of heat and hydrogen gas.

The catalyst is mixed with tap water in a proportion of about 225 g. per liter of water. In other words, the sodium hydroxide content of the catalytic solution is preferably about 18% by weight. The catalyst is not

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The aluminum used in the reaction comprises aluminum foil, electrical wire, beverage cans and other similar aluminum waste. The intensity of the reaction depends upon the surface of contact between the aluminum and water. Aluminum foil for example reacts faster than a

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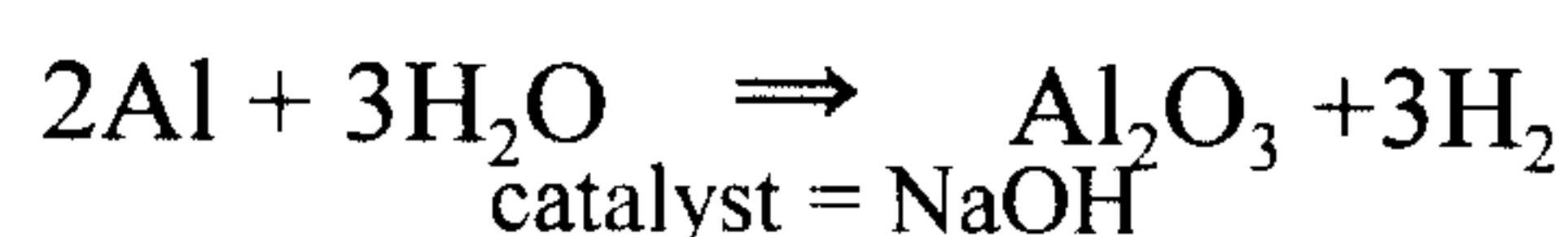
A series of eight experiments was carried out to measure the volume of hydrogen gas produced in a typical reaction. In these experiments, aluminum foil from Reynolds Aluminum Company of Canada was loosely

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The volume of water displaced by the gas produced was measured and corrected to a gas volume at standard temperature and pressure (STP). Atmospheric pressure on that day was obtained from a local weather office. The corrected volume of gas produced was compared to the theoretical quantity of hydrogen gas, which would be obtained according to the equation,

20

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These experiments were carried out at a room temperature of 21 °C and an atmospheric pressure of 758 mm of Hg. In all cases the reaction started in few seconds and continued for few minutes, until depletion of the aluminum foil. It was noticed that a typical reaction with less than 5 grams of loosely crumpled aluminum foil, is complete in less than 5 minutes. The results of these experiments are shown in **Table 1** below.

Table 1: Hydrogen Gas Production from Aluminum Foil

Exp. (#)	Al (g.)	H ₂ (l)	H ₂ (l) (STP)	H ₂ (l) Theoretical	Yield (%)	Deviation (+/- %)
1	2.08	2.94	2.71	2.59	104	2.6
2	2.03	2.85	2.62	2.53	104	2.6
3	2.21	3.05	2.81	2.75	102	2.5
4	2.16	2.9	2.67	2.69	99	2.6
5	2.2	3.04	2.8	2.74	102	2.5
6	2.21	3.04	2.8	2.76	102	2.5
7	0.73	1.03	0.94	0.91	103	2.4
8	0.83	1.15	1.05	1.03	102	2.2
Ave.					102	2.47

The results from **Table 1** show that the reaction is reproducible and produces stoichiometric quantities of hydrogen gas. The 102% average yield of hydrogen gas is considered to be within the measurement uncertainty; however, there are at least two factors which might have contributed to a slightly higher hydrogen yield. Firstly, the volume of gas produced was corrected to STP. It is possible that the exhausted fume hood in which the experiments were carried out could have lowered the reaction pressure below the atmospheric pressure of 758 mm of Hg. This would have increased the observed value for the volume of gas produced.

An exhaust bench typically runs at 1 inch or 2 inches of water pressure. At a maximum, this could have increased the measured volume by about 0.5%. Secondly, the water used was tap water in all cases, in which dissolved air may have been present. If any of this air had been released in the presence of the warm hydrogen gas, this would have increased the volume of gas measured. This would have affected the results by less than 1%. Since the results are within the measurement error, and quantification of these two sources of error would not significantly affect the results, no further experiment was carried out in this area

The procedure used in the above experiments was repeated, with the exception that the tube leading from the top of the reaction bottle was connected to a gas sampling bag. Two samples of gas were obtained and analysed. The results are presented in **Table 2**.

Table 2: Gas Analysis

Sample	Hydrogen Concentration	Oxygen & Nitrogen
1	92 %	balance
2	98 %	balance

Table 2 shows that the purity of the hydrogen collected in the second sample was 98%. This is close to what was theoretically expected. The lower 92% concentration observed in the first sample was probably due to the fact the system was not completely purged with hydrogen before the sample was taken. By the time the second sample was taken, most of the air had been purged from the tube and the reaction bottle.

The procedure used in the first mentioned experiments was repeated except that the reaction bottle was placed in a water bath before the

aluminum was added to the water, and the hydrogen produced was bubbled through the bath water. The temperature of the bath and the catalytic solution were measured before and after the reaction, and at about four minutes after the reaction was completed.

- 5 The water equivalent of the plastic containers for absorbing heat and their specific heat were determined experimentally by adding a known quantity of hot water to the reaction system at room temperature and then calculating the heat transfer based on the final temperature.

- 10 The quantity of heat produced by the reaction was determined and compared with the theoretical values. The results are shown in **Table 3**.

Table 3: Heat of the Reaction

Readings	Temp. °C Reactor Start	Temp. °C Bath Start	Temp. °C Reactor Finish	Temp. °C Bath Finish	Time
1	21.1	20.2	45.5	24.4	5.29
2	21.1	20.2	38.3	25.3	5.33

- 15 Table 3: Continued

Readings	Al (g)	Heats of Formation Al ₂ O ₃ kcal/mole	Heat of Formation H ₂ O kcal/mole	Heat Output Actual kcal	Heat Output Theoretical kcal	Efficiency (%)
1	9.52	-400.5	-68.3	33.3	34.5	96
2	9.52	-400.5	-68.3	32.5	34.5	94

- 20 The results in **Table 3** show that the observed heat released in the production of hydrogen was 96% of the theoretical value. The 94% value from the second reading can be attributed to the heat lost to the surroundings during the time that lapsed between the readings.

The reaction has a net maximum heat production during hydrogen generation of 195.6 kCal/mole. A further 204.9 kCal/mole will be released if the hydrogen is burned with oxygen. Stated another way, 51% of the reaction energy is used to form hydrogen gas and 49% goes into the production of heat.

Having explained the preferred method for producing hydrogen gas, the following disclosure and drawings describe a preferred apparatus for carrying out the method.

Referring firstly to **FIGS. 1 and 2**, an energy production apparatus according to the preferred embodiment of the present invention is illustrated therein. The energy production apparatus, also referred to herein as the hydrogen generator **20** is illustrated in these figures in its entirety. The hydrogen generator **20** uses water and aluminum particles as fuel, and sodium hydroxide (NaOH) as a catalyst and a surface conditioner to reduce the formation of oxide layers on the aluminum particles. The sodium hydroxide may be mixed or otherwise closely associated with the aluminum particles in a sufficient amount to ensure complete reaction of the aluminum particles with water in an energy production period. Further discussion on the incorporation of sodium hydroxide with the aluminum particles will be provided later, especially when making reference to **FIGS. 10-13**.

The hydrogen generator **20** comprises firstly a receptacle **22** having a first closable fill opening **24**, a sight glass **26** for monitoring the level of water therein, and a second closable larger opening **28** in a central upper region thereof. The receptacle **22** also preferably has a cleanout bung **30** through its bottom surface to facilitate the periodic removal of the reaction

byproducts such as alumina.

In use, the receptacle **22** is filled with water **32**, to a level of between half and three-quarter of its capacity. A fuel cartridge **34** hanging from a vertical tube **36** is immersed into the water **32** for causing a chemical reaction to occur with the water, and for producing heat and hydrogen gas.

The fuel cartridge **34** is supported in a perforated basket **38** affixed to the vertical tube **36**. The vertical tube **36** is connected to a gas handling manifold **40** mounted above the receptacle **22**, and has a series of holes **42** therein for admitting the hydrogen gas into the gas handling manifold **40**.

An annular cap **44** is also provided for mounting over the upper central opening **28** of the receptacle. Several clasps **46** are provided around the annular cap **44** for securing the annular cap **44** in a sealing manner to the upper central opening **28**. It will be appreciated that the upper central opening **28** has a dimension to accommodate the insertion of the fuel cartridge **34** and the basket **38** inside the receptacle **22**. It will also be appreciated that the clasps **46** may be replaced by other closure means for quickly and easily removing the annular cap **44**, for replacing a spent fuel cartridge for example.

Upon the annular cap **44**, there is provided a bellows **48** having an interior region communicating with the receptacle **22** such that the expansion and retraction of the bellows are relative to the pressure inside the receptacle. Atop the bellows **48**, there is provided a timer mechanism **50**, the operation of which will be described later. In the hydrogen generator according to the preferred embodiment **20**, the gas handling

manifold **40** is affixed to the upper portion of the vertical tube **36** above the timer mechanism **50**.

5 The annular cap **44**, the bellows **48**, the timer mechanism **50** and the gas handling manifold **40** define with the receptacle **22** a closable space for containing and controlling the hydrogen gas being generated inside the receptacle **22**.

10 A burner plate **60** is mounted over the gas handling manifold **40**. The gas handling manifold **40** has conduit means communicating with the burner plate **60**. A series of orifices are provided in the burner plate **60** to allow the burning of hydrogen gas for cooking food for example in a similar manner as is known of gas stoves. Although the illustrations show a side view of the burner plate **60** it will be appreciated that the burner plate **60** is preferably a circular plate similar to those mounted on common gas stoves.

15 The gas handling manifold **40** also has a selector valve **62** and a gas outlet fitting **64** communicating with the selector valve **62**. The selector valve **62** is operable for selectively directing the hydrogen gas to the burner plate **60** or to the outlet fitting **64**.

20 A flexible hose for example, as represented by dash lines **66**, can be connected to the outlet fitting **64** and to a gas appliance such as a lantern **68** to conveniently use the hydrogen gas, or to a gas reservoir **70** for accumulating the hydrogen gas for later use. The appliance may have a water filter **72** thereon if needed or a check valve **74** to prevent any
25 backflow of gas into the receptacle **22** at the end of an energy production period.

With reference to **FIGS. 2** and **14**, it will be appreciated that a typical energy production period is known to have a heating phase 'A' during which the temperature inside the receptacle **22** rises; an active phase 'B' during which the temperature inside the receptacle **22** is preferably kept at around 85° C, and a cooling phase 'C' during which the reaction gradually stops. At an operation temperature 'T' during the active phase 'B' of about 85° C, the reaction has been found to be self-sustained and the hydrogen gas produced contained minimum water vapours.

The heating phase 'A' can be shortened by introducing a fuel pellet **80** inside the receptacle **22**, through the fill opening **24**. The fuel pellet **80** preferably contains very fine aluminum particles such as saw dust and filings for examples, compressed with waste paper bits that are impregnated with sodium hydroxide in a dry form. The small aluminum particles of the pellet **80** are known to be highly reactive with water to generate a burst of heat which causes the water temperature to approach the ideal temperature 'T' quickly, and to accelerate a reaction of the water with the larger fuel cartridge **34**. Another fuel pellet **80** may also be introduced in the receptacle during the cooling phase 'C' to prolong the duration of an energy production period.

For example purposes, a fuel cartridge **34** having a volume of about one liter, that is about 500 ml of aluminum and about 500 ml of paper filler material impregnated with sodium hydroxide in a dry form, immersed in 10 liters of water is believed to be sufficient for producing heat and maintaining a reaction for about two hours, in which the active phase is about one hour, and the heating and cooling phases are about one-half hour each. It is believed that the amount of hydrogen gas produced during the active phase 'B' is sufficient for cooking food on the burner plate **60**.

Referring now to **FIG. 3**, the operation of the bellows **48** is illustrated therein. When the reaction enters its active phase, the heat and pressure generated inside the receptacle **22** rise. The increase in pressure inside the receptacle **22** causes the bellows **48** to expand upward as
5 illustrated in **FIG. 3**.

Because the basket **38** and the vertical tube **36** are supported to the gas handling manifold **40**, and because the gas handling manifold **40** is affixed to the movable portion of the bellows **48**, the expansion of the bellows **34** causes the fuel cartridge **34** to be lifted toward an upper region
10 of the receptacle **22**, and by the same doing, causes the water level to fall in the receptacle **22**. The contact surface between water and the fuel cartridge **34** is thereby greatly reduced. The reaction is slowed down and the pressure and temperature inside the receptacle **22** are consequently also reduced. As temperature and pressure inside the receptacle **22** are reduced,
15 the bellows **48** collapses to re-immersed the fuel cartridge **34** and to resume the active reaction phase.

Given the structure of the energy production apparatus **20** according to the preferred embodiment, it is believed possible to calibrate the characteristics of the bellows **48** for use with a specific size of receptacle
20 **22** and a specific size of fuel cartridge **34**, to precisely control the pressure and temperature of a reaction, such that the apparatus **20** will be practical and safe for use by the general public.

With reference to **FIGS. 4, 8 and 9**, the functions of the timer mechanism **50** of the hydrogen generator **20** will be explained in details.
25 The timer mechanism **50** is provided for further improving the safety of the hydrogen generator **20**. The timer mechanism **50** is used for lifting the fuel

cartridge **34** above the water **32** after a set time period, even when the bellows **48** remains in a collapsed mode. The reaction inside the receptacle **22** can thereby be manually stopped or caused to terminate at a set time period by adjusting a knob **90** relative to a dial **92**.

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The preferred timer mechanism **50** comprises a coil spring **94** mounted over the vertical tube **36** and an annular spring-abutment plate **96** affixed to the vertical tube **36** above the spring **94** for retaining the vertical tube **36** at a fixed position relative to the upper end of the spring **94**.

10

The spring **94** is set in a cylindrical pocket **98** extending downward through the timer mechanism **50**. The depth of the pocket **98** is sufficient to accommodate the spring **94** in a compressed form when the timer mechanism is in a latched mode. A seal **100** is affixed to the bottom portion of the pocket **98** around the vertical tube **36**, for allowing a sliding movement of the vertical tube **36** through the timer mechanism **50**, under the action of the spring **94**, and for preventing hydrogen gas from leaking out of the bellows **48**.

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One or more latch tabs **102** are movably connected to the timer mechanism **50** and are linked to the operation of the selector knob **90**. When the burner plate **60** is pushed down to immerse the fuel cartridge **34** in water, the latch tabs **102** engage with the annular spring-abutment plate **96** to keep the spring **94** in a compressed state inside the cylindrical pocket **98**.

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The linkages, the clockwork and other components mounted inside the timer mechanism **50** have not been illustrated herein for being common to those knowledgeable in latches and locks. In the preferred embodiment,

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however, the clockwork is a mechanical device not requiring electric power. Also in the preferred embodiment, the latched tabs **102** are in a latching position when the timer knob **90** is set at any time value, and are in an unlatching position when the knob **90** is set at or reaches zero (0) time on the dial **92**.

Referring now to **FIG. 5**, the structure of the burner plate **60** and of the gas handling manifold **40** are explained therein in greater details. The outline of the gas handling manifold **40** is shown in dash lines to simplify the illustration. The gas handling manifold **40** comprises a first set of conduits **110** extending from the vertical tube **36** to the selector valve **62**, to a pressure relief valve **112**, and to a flow control valve **114**; a second set of conduits **116** extending from the selector valve **62** to the burner plate **60**; and a third conduit **118** extending from the selector valve **62** to the outlet fitting **64**.

The burner plate **60** has a plurality of gas orifices **120** therein, and each gas orifice is preferably surrounded by one or more air injection holes **122** to admit oxygen around the gas orifice **120** during the burning of hydrogen gas.

In the preferred embodiment, a minimum amount of hydrogen gas is always directed to the gas orifices **120** to be burnt. The burning of this minimum amount of gas provides a visual indication of the operation of the apparatus **20**, and prevents any accumulation of hydrogen gas in the room in which the apparatus is being used. For this purpose, a flow control valve **114** is provided in the gas handling manifold **40**, and has a fourth conduit **124** bypassing the selector valve **62**. Therefore, when the selector valve **62** is set to direct the hydrogen gas to the outlet fitting **64**,

a minimum amount of gas is still allowed through the flow control valve **114** and to the gas orifices **120** of the burner plate **60**.

5 The flow control valve **114** is preferably an adjustable type such that it can be opened fully to bypass both the selector valve **62** and the pressure relief valve **112**, to obtain a larger flame **126** at the center of the burner plate **60** if needed.

10 The pressure relief valve **112** is provided to further improve the safety of the apparatus, as will be understood from the following description. The pressure relief valve **112** monitors the pressure inside the vertical tube **36** and releases a pressure over an unsafe level, to a whistle **128** which has an outlet opening positioned near one of the gas orifices **120**. The gas flowing from the whistle **128** may thereby be readily ignited by the flame above that orifice **120**, to provide a visual indication of an abnormal operation of the apparatus. The sound of the whistle **128** is yet
15 another sign to alert a user of an over pressure inside the receptacle **22**, and to urge that user to set the knob **90** to zero time to cause the timer mechanism **50** to raise the fuel cartridge **34** out of the water.

20 As illustrated in **FIG. 6**, the burner plate **60** preferably has a pair of handles **130** affixed thereto to manipulate the upper portion of the apparatus **20** when the clasps **46** are released and the basket **38** is lifted out of the receptacle **22**.

25 In the preferred apparatus, a gas filter **132** may also be installed over the gas admitting holes **42**, for preventing any accumulation of reaction byproducts inside the vertical tube **36**. In other embodiments, the vertical tube **36** may be filled with an appropriate granular filtering

medium for example for preventing reaction byproducts from reaching the gas handling manifold **40**.

As will be appreciated, a pressure gauge **134**, a temperature gauge or both, may also be provided on the annular cap **44** or at another convenient location allowing a communication with the receptacle **22**, for visually monitoring the development of a reaction occurring inside the apparatus.

In **FIG. 7**, the burner plate **60** is shown supporting a heat storage device **140**, for storing heat during the operation of the apparatus **20**. The heat storage device **140** is used for prolonging the beneficial effect of an energy production period when the apparatus **20** is used to heat a camp in the wilderness, or a household during a power outage period for example. The preferred heat storage device **140** comprises a copper plate **142**, supported on legs **144**, above the burner plate **60**, and a perforated dome-shape enclosure **146** enclosing one or more rocks **148** laid over the upper portion of the copper plate **142**. The heat storage device **140** is removable from the burner plate **60** and is preferably used whenever the burner plate **60** is not used for cooking food. Further, the receptacle **22** is preferably made of steel or similar heat conductive material for radiating heat during the entire energy production period.

In the preferred embodiment, the inside diameter of the bellows **48** is sufficiently large, 15-25 cm for example, and the spring **94** is calibrated such that the weight of the heat storage device **140** or the weight of a common cooking pot (not shown) which may be set on the burner plate **60** does not significantly affect the operation of the bellows **48** or of the timer mechanism **50**.

Referring now to **FIGS. 10-13**, several arrangements are proposed for preparing the fuel elements required for use in the hydrogen generator **20** according to the preferred embodiment. The fuel bundle **34** is preferably prepared by overlaying a thin strip of aluminum **150** over a
5 sheet of embossed paper **152** impregnated with sodium hydroxide in a dry form. The aluminum sheet and the paper layer are coiled together to form a cylindrical shape. The preferred cartridge **34** is loosely coiled such that water may be readily absorbed between the layers of the cartridge. The advantage of a loosely coiled cartridge **34** is that the water is allowed to
10 seep into the entire cartridge at once to create an intense reaction, and reduce the duration of the heating phase 'A' of the reaction as illustrated in **FIG. 14**.

Another advantage of the cartridge **34** as described above is that when the layers of the coil are set vertically, the hydrogen gas generated
15 between the layers rises up and creates a vacuum between the layers at the lower end of the cartridge **34** to admit more water from the lower end of the cartridge. This phenomenon is advantageous for wetting the aluminum strip quickly, entirely and continuously.

The preferred fuel cartridge **34** is packaged in a sealed envelope that
20 has an indication as to its duration, potential heat energy and volume of hydrogen gas to be produced by it.

Although the fuel cartridge **34** may be better manufactured with virgin material, it will be appreciated that there are numerous economical and environmental advantages in the manufacturing of fuel elements from
25 waste materials. Accordingly, another preferred fuel unit **160** having a

loose content in a bag-like envelope is illustrated in **FIG. 11**. The envelope **162** is water-permeable, and the loose content comprises aluminum turning, aluminum saw dust and filings, aluminum shreds and other aluminum waste particles **164** as normally found in a metal working shop, or as available from scrap metal vendors.

It is also possible to use aluminum shreds from domestic waste containers. When the waste aluminum is obtained by shredding food or drink containers for example, the waste material is preferably pre-treated to at least partly remove a protective coating on this aluminum material.

The loose content of the fuel unit **160** also comprises waste paper bits **166** impregnated with sodium hydroxide and dried. The paper bits **166** are made of waste newsprint or similar recyclable paper waste. The paper bits **166** preferably have sizes and quantities similar to the aluminum particles, and are mixed with the aluminum particles **164**. The presence of the paper bits **166** prevents the fusion of the aluminum particles **164** together and ensures a continuous absorption of water throughout the loose content of the fuel unit **160**. The fuel unit **160** is also preferably manufactured and labelled as to indicate its expected energy production period.

The fuel pellet **80** as illustrated in **FIGS. 2** and **12**, and as previously described contains very fine aluminum particles such as saw dust and filings for example, to provide a better water contact and a more intense reaction. One or two fuel pellets **80** are preferably packaged in a sealed envelope and distributed as reaction accelerators with each fuel cartridge **34**, or with each fuel unit **160** sold.

A third preferred arrangement for a fuel element usable in the apparatus **20** according to the preferred embodiment is illustrated in **FIG. 13**. The fuel measure **170** is preferably comprised of a perforated container **172** filled with aluminum waste **164** and paper bits **166** as previously described.

When the energy production apparatus according to the preferred embodiment **20** is used by someone having access to aluminum waste material, and who does not want to depend on purchased fuel elements, the fuel measure **170** described herein is recommended and is preferably used with a nominal quantity of sodium hydroxide **174** set over the fuel measure **170**. The sodium hydroxide **174** may be compressed into a tablet form as illustrated for easy handling and storage. It may be used in a powder form contained in a water-permeable sachet (not shown), or may be kept in a sealed container and sprinkled generously over the water **32** before introducing the fuel measure **170** into the water **32**.

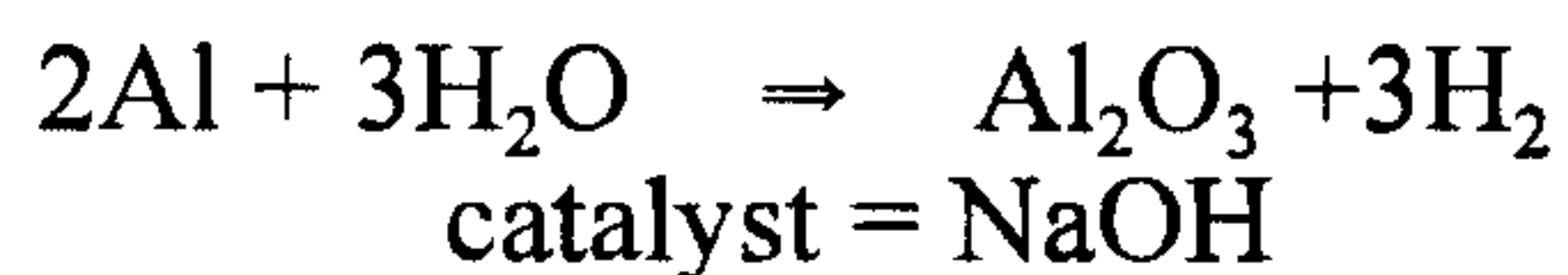
As to the manner of manufacture of the preferred apparatus, the same should be apparent from the above description and accompanying drawings, and accordingly further discussion relative to this aspect would be considered repetitious and is not provided.

While one embodiment of a method and of one apparatus have been described herein above, it will be appreciated by those skilled in the art that various modifications, alternate materials, compositions and equivalents may be employed without departing from the true spirit and scope of the invention. Therefore, the above description and illustrations should not be construed as limiting the scope of the invention which is defined by the appended claims.

CLAIMS

We claim:

1. A process for producing hydrogen gas, comprising the steps of
 providing an aqueous solution containing about 4.5 M NaOH;
 reacting aluminum with water in said solution; and
 limiting a temperature of said solution to about 45°C,
 such that a reaction of said aluminum with said water is a catalytic
 reaction with said NaOH acting as a catalyst, and is carried out
 according to the following formula:



2. A process for producing heat, light and hydrogen gas, comprising
 the steps of:
 providing an expandable receptacle;
 partly filling said expandable receptacle with water;
 introducing an aluminum element and sodium hydroxide as catalyst
 in said water;
 partly sealing said expandable receptacle;
 reacting said aluminum element with said water in the presence of
 said sodium hydroxide catalyst, and
 expanding and contracting said expandable receptacle and
 emerging and immersing said aluminum element and said
 sodium hydroxide catalyst out of and in said water, in
 response to more or less pressure inside said expandable
 receptacle respectively;
 such that an intensity of a reaction between said water and said
 aluminum element is controllable in response to a pressure
 generated inside said expandable receptacle by said reaction.

3. The process as claimed in **claim 2**, wherein said step of reacting said aluminum element with said water is initiated at a temperature of about 21⁰C.
4. The process as claimed in **claim 2**, wherein said aluminum element is a coiled strip of aluminum having several layers set vertically in said water, and said process further comprises the step of absorbing water by vacuum through said layers.
5. An apparatus for producing heat, light and hydrogen gas, comprising;
an expandable receptacle having an upper end, a central portion,
bellows mounted to said upper end and a sealing means
connected to said bellows for partly sealing said receptacle;
said bellows having a fixed portion connected to said upper end of
said expandable receptacle and a movable portion movable
toward and away from said upper end of said expandable
receptacle;
a fuel basket mounted inside said expandable receptacle in said
central portion of said expandable receptacle, said fuel
basket being connected to said movable portion of said
bellows; and
said expandable receptacle being adapted to contain water in said
central portion thereof, and said fuel basket being adapted to
contain an aluminum element;
such that said fuel basket is movable toward and away from said
central portion of said expandable receptacle for immersing
and lifting an aluminum element in and out of said water
receptacle, to control a chemical reaction between said
aluminum element and said water.

6. The apparatus as claimed in **claim 5**, further comprising a timer mechanism and means responsive to said timer mechanism for timely raising said fuel element in said central portion.
7. The apparatus as claimed in **claim 5**, further comprising a burner plate mounted over said upper end.
8. The apparatus as claimed in **claim 7**, further comprising a gas conduit, a gas outlet fitting and a selector valve joining said gas conduit, said gas outlet fitting and said burner plate.
9. The apparatus as claimed in **claim 5**, wherein said fuel element comprises a coiled strip of aluminum having several layers set vertically in said central portion.

Mario D. Theriault, P.Eng.

“Patent Agent of the Applicants”

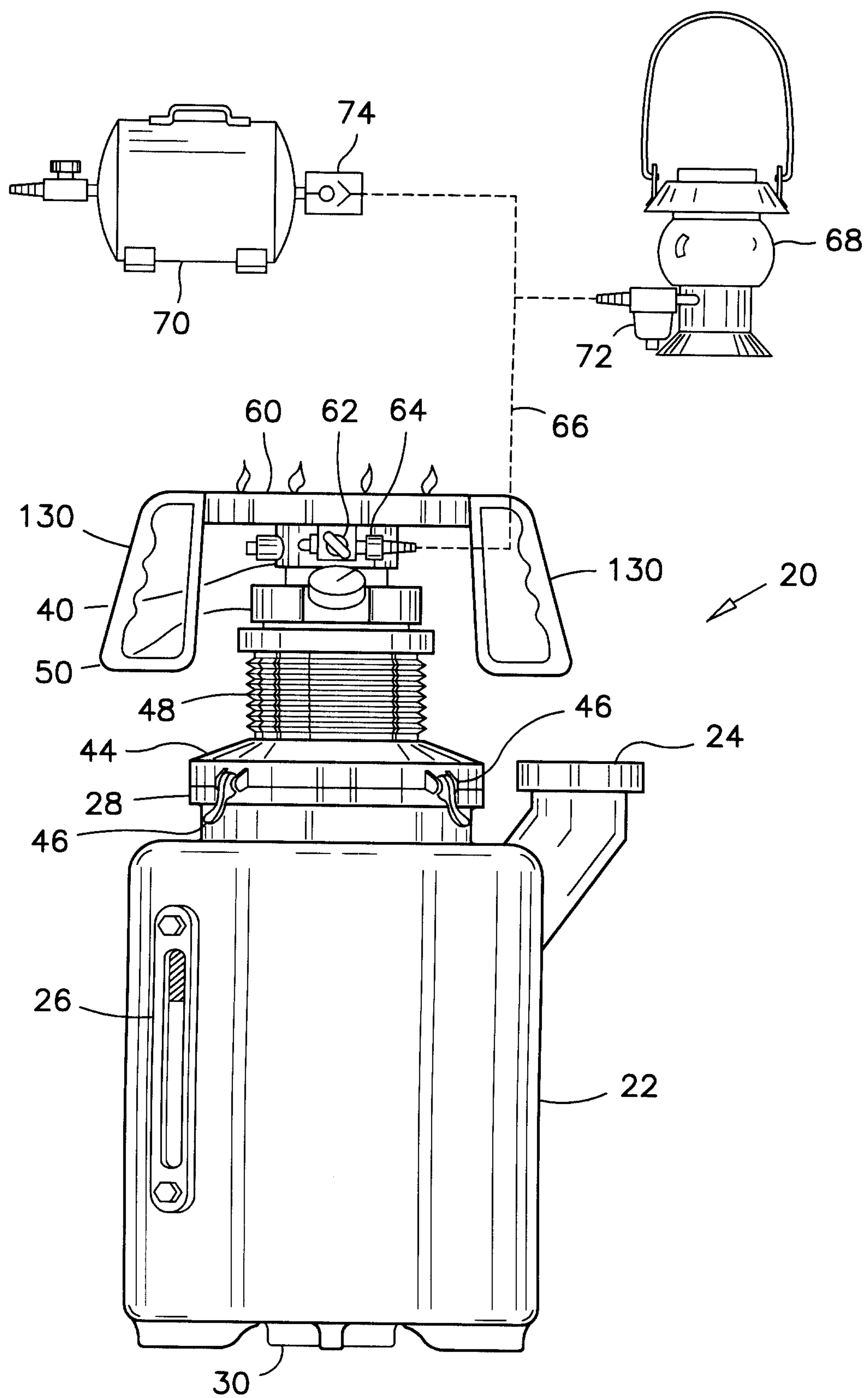


FIG. 1

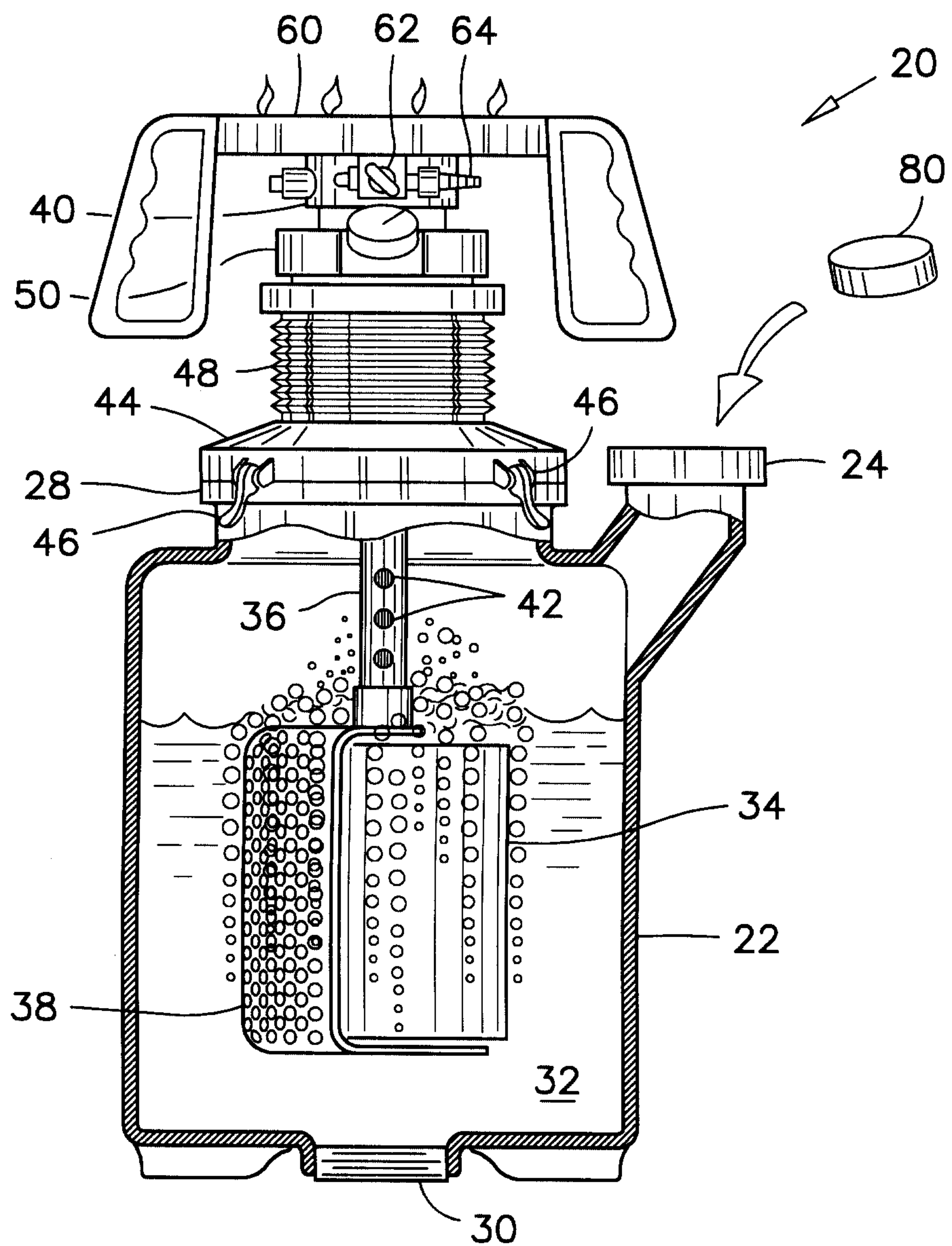


FIG. 2

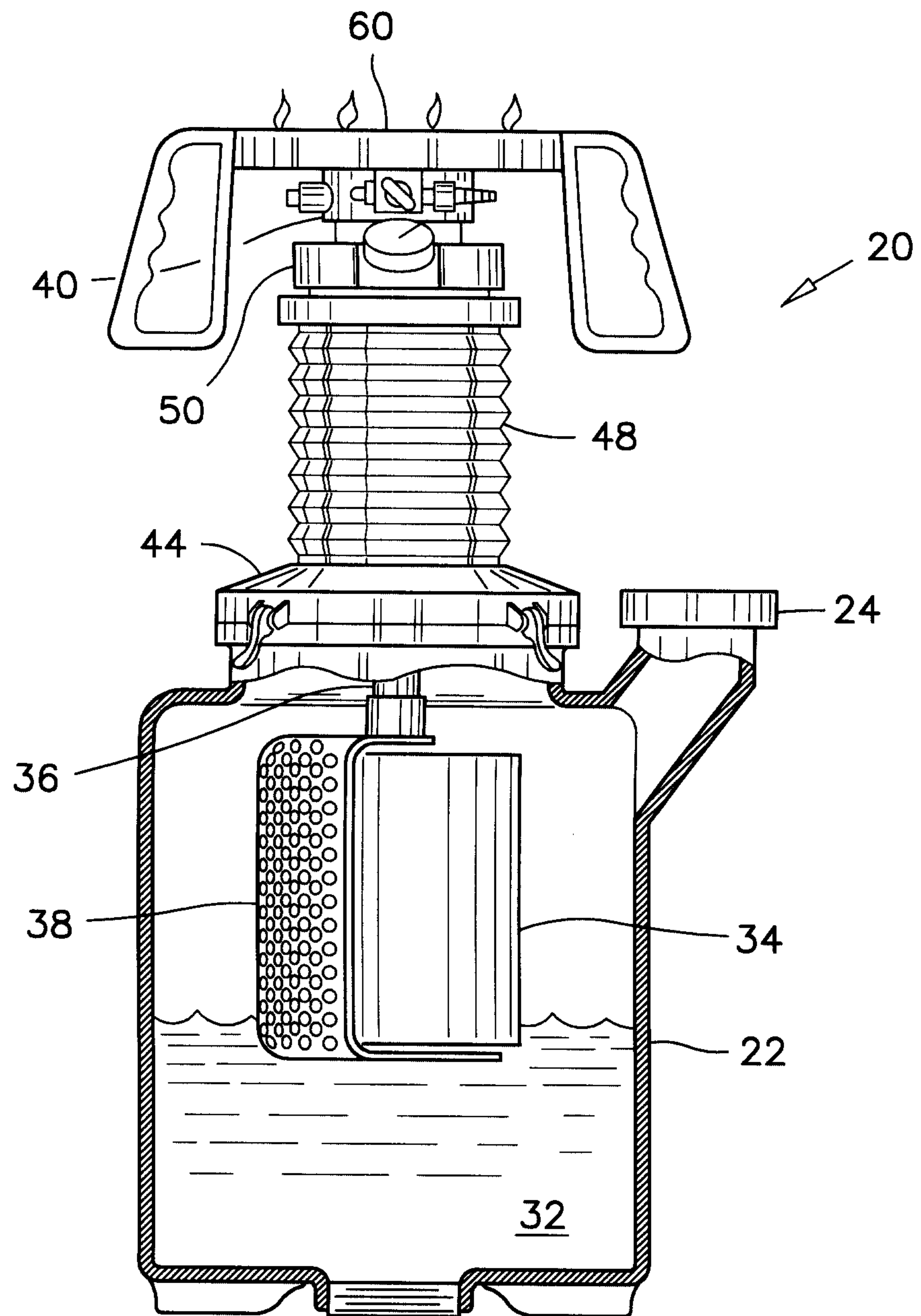


FIG. 3

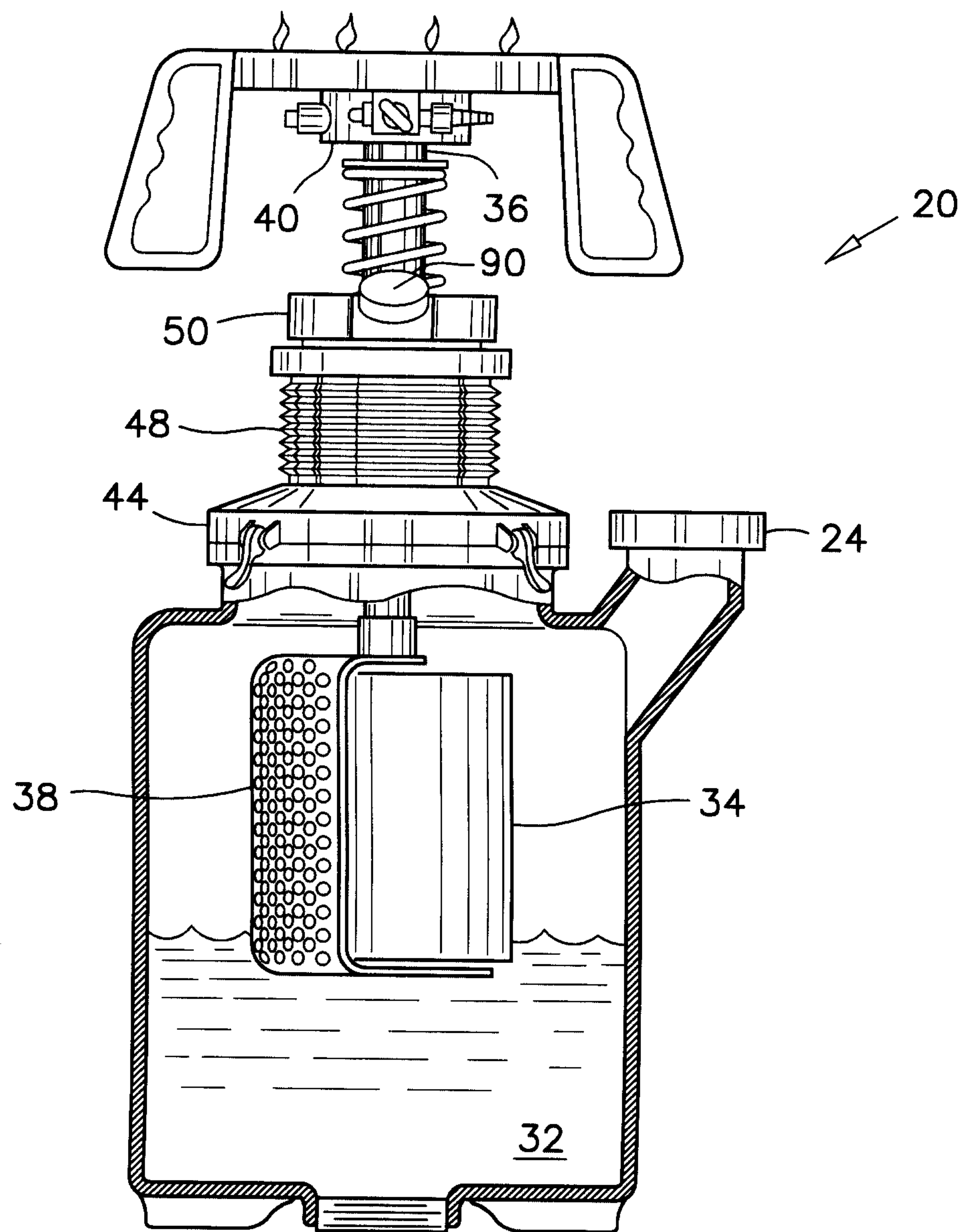


FIG. 4

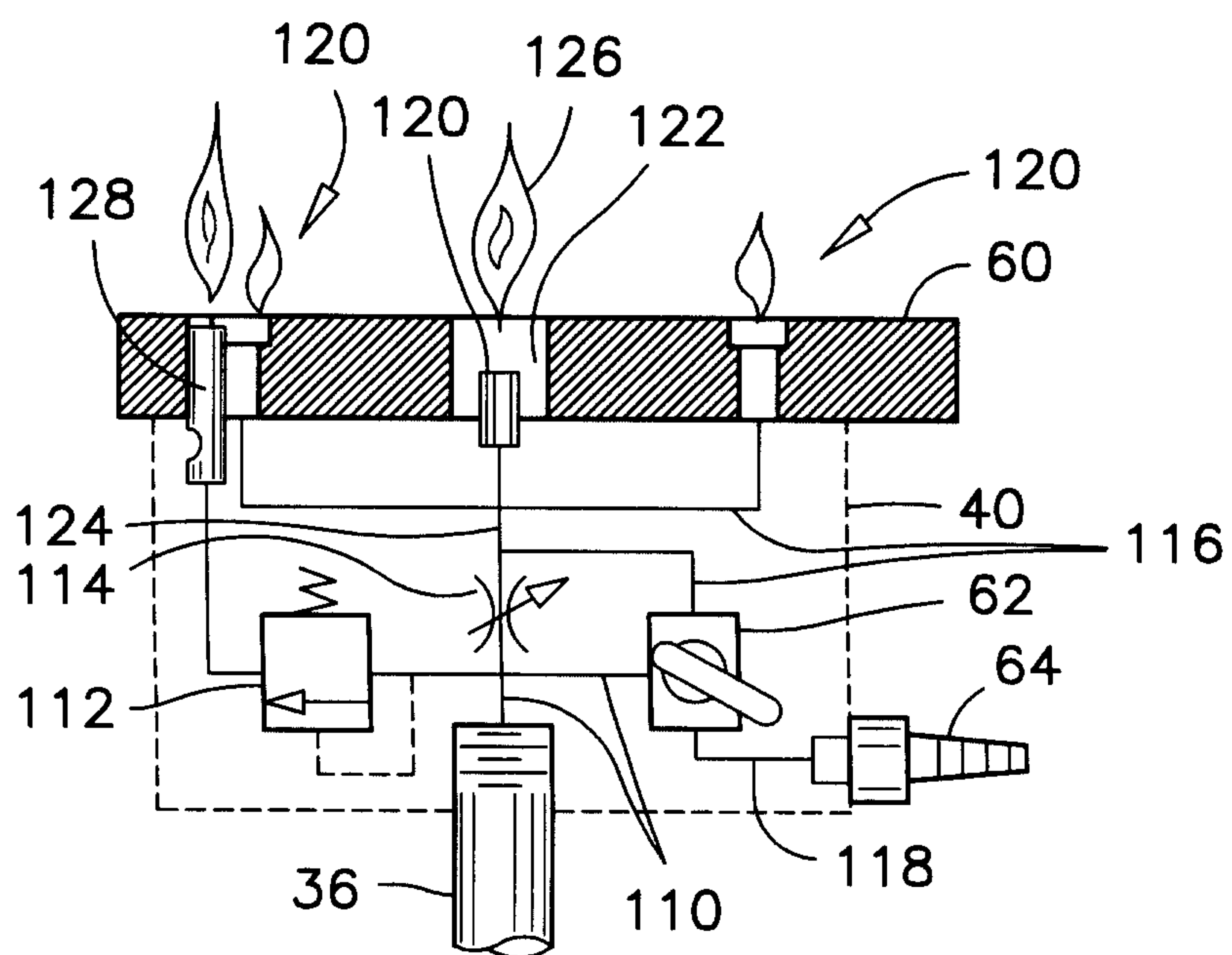


FIG. 5

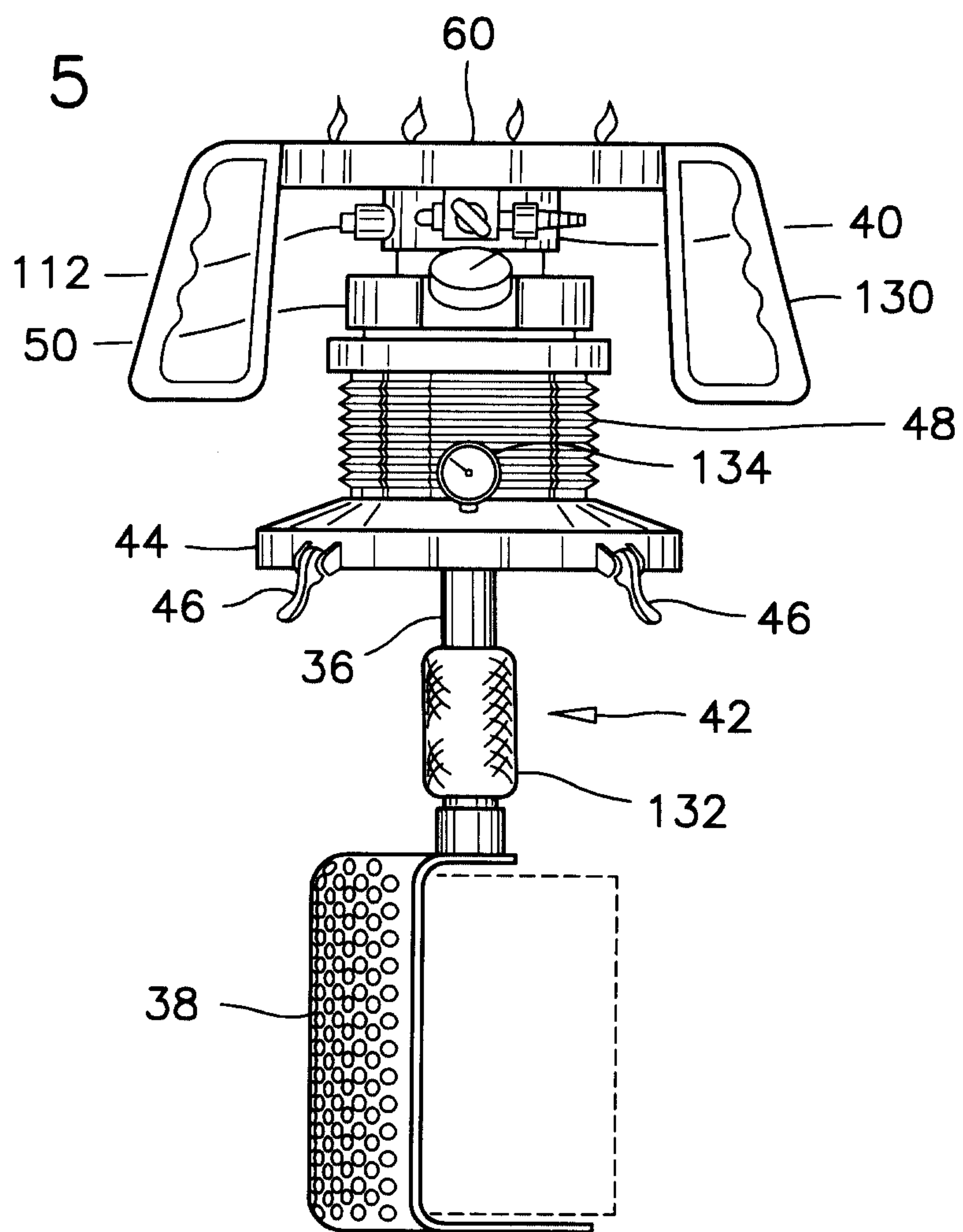
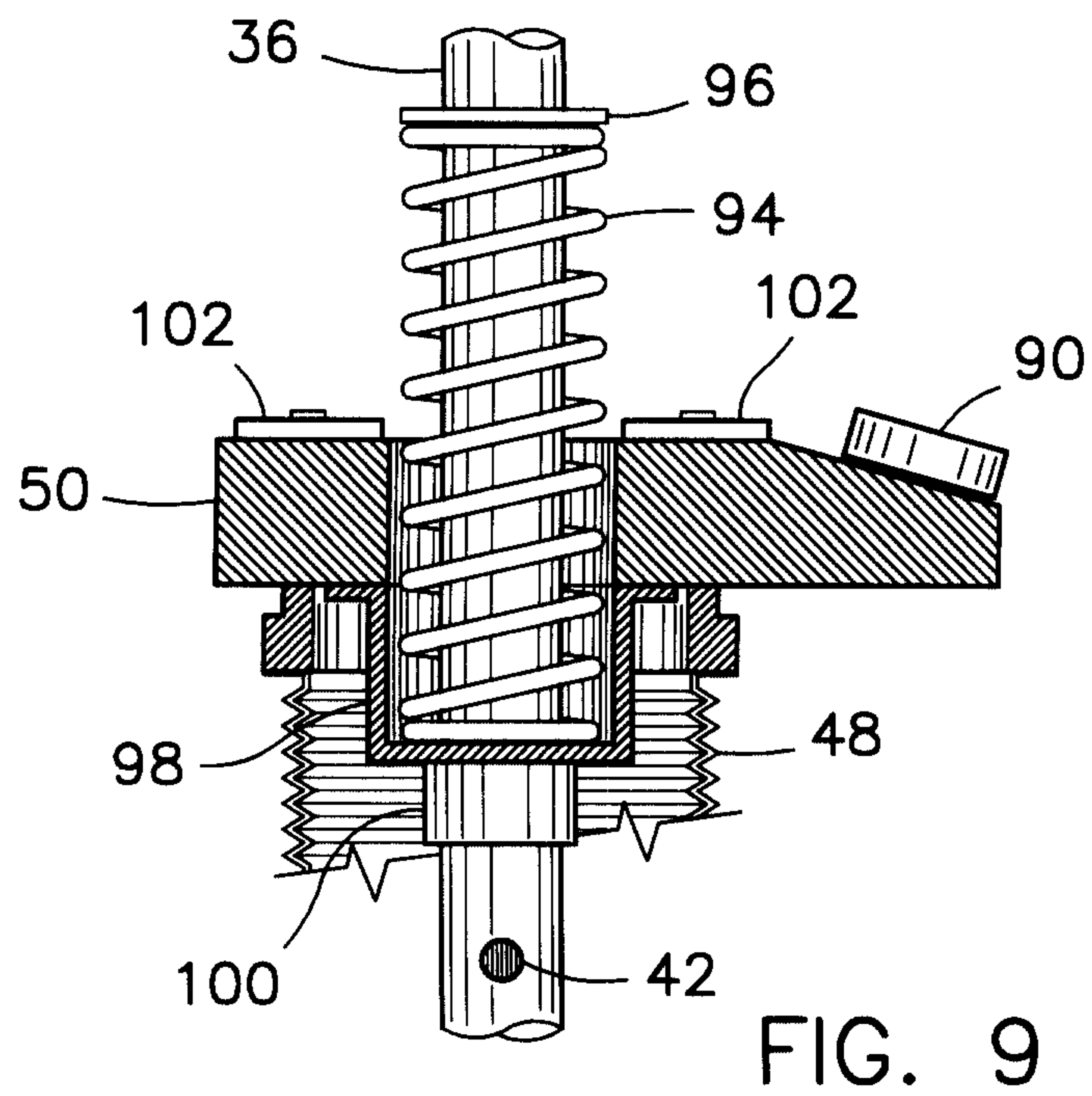
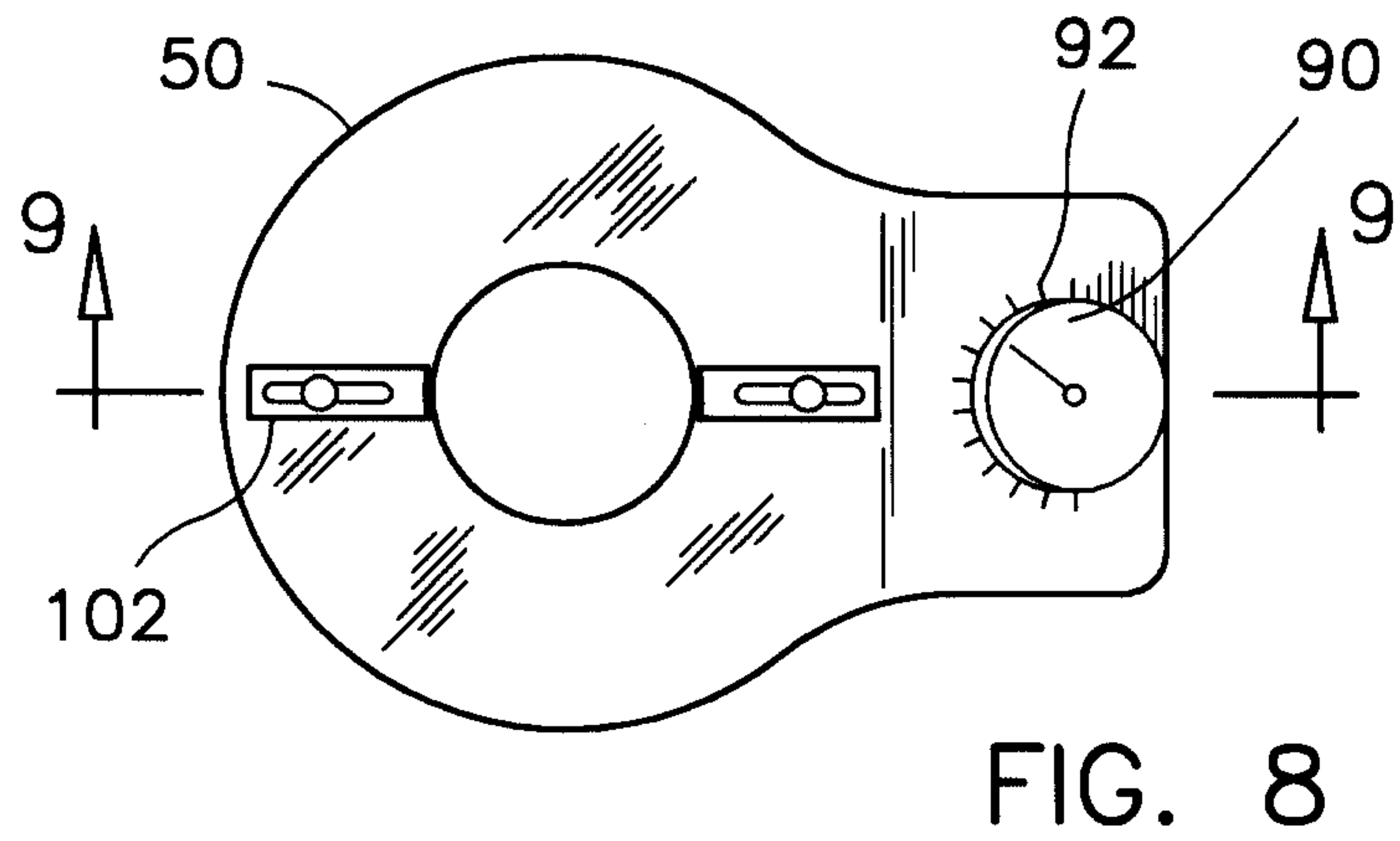
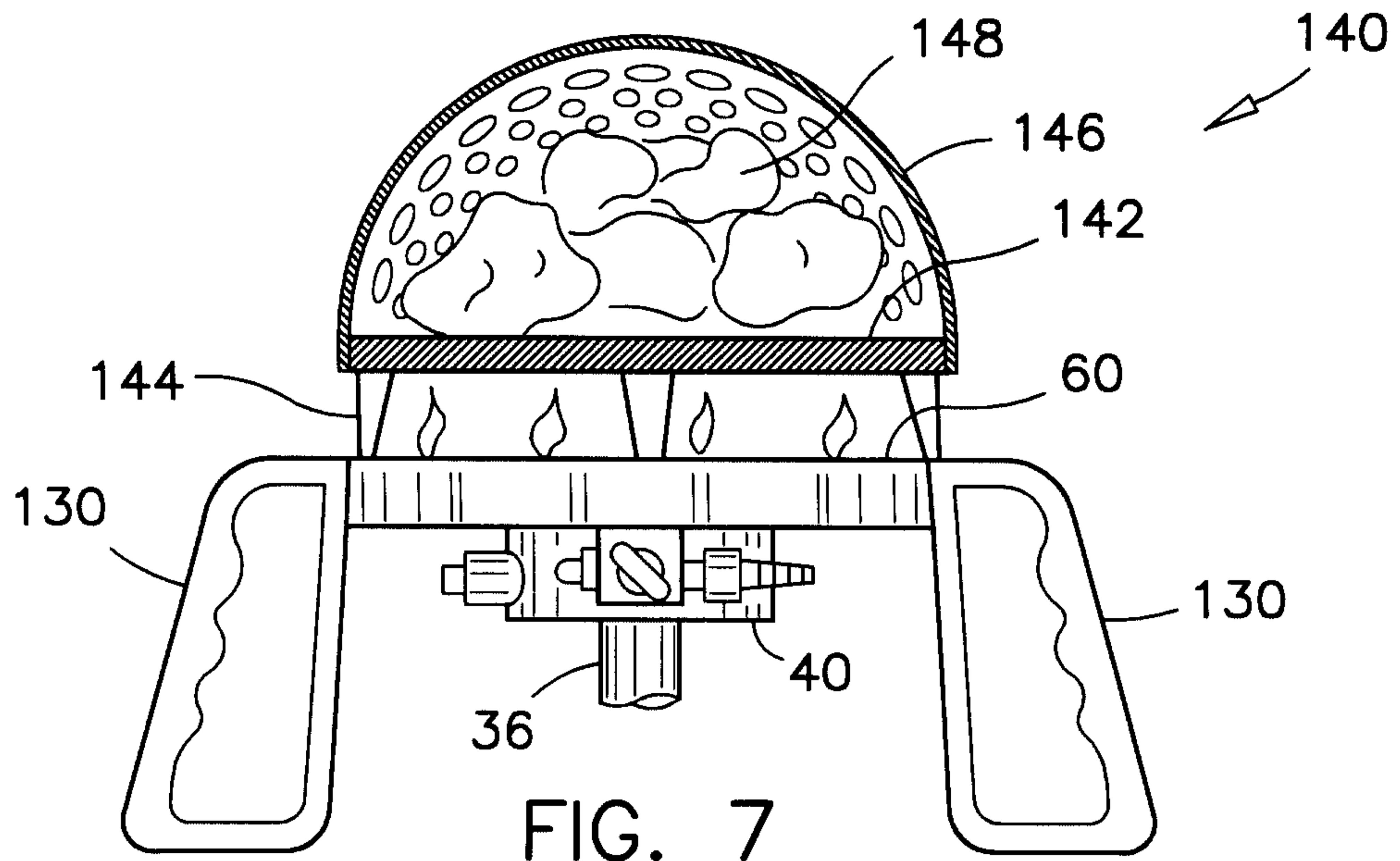


FIG. 6



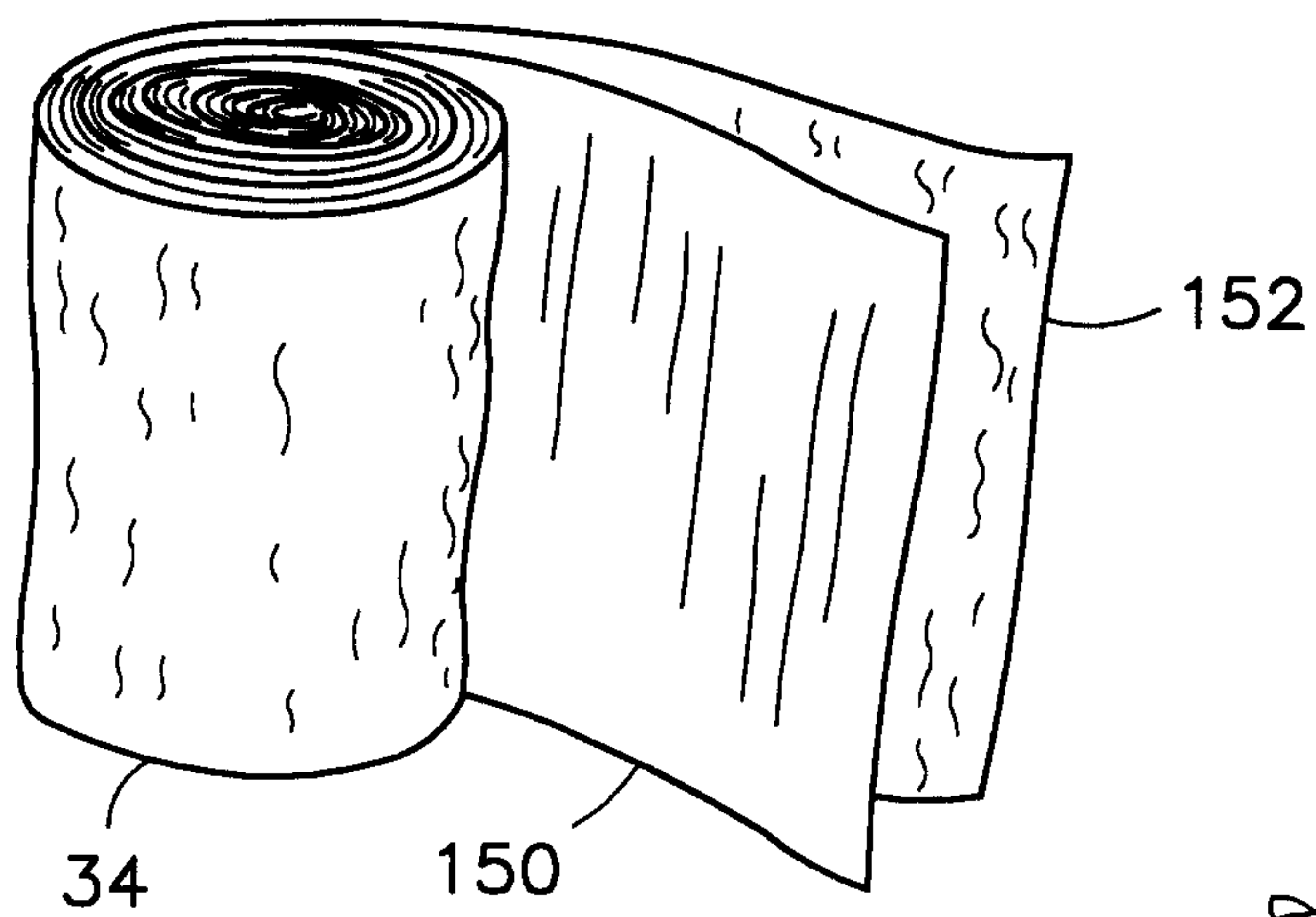


FIG. 10

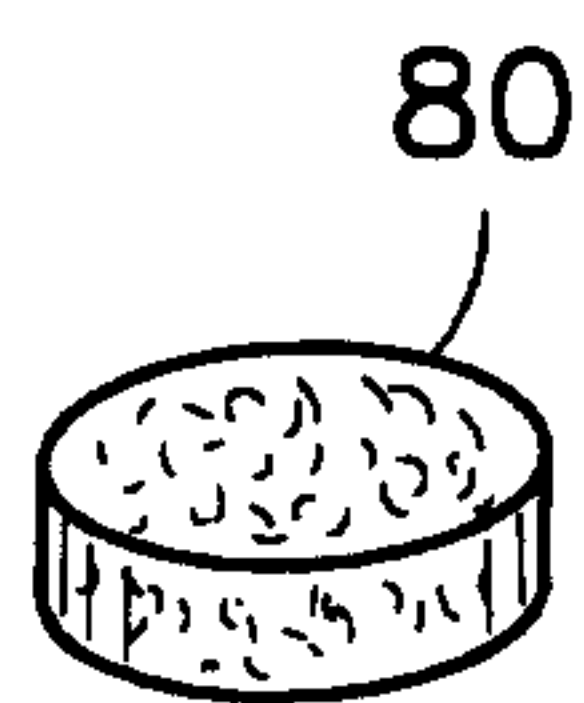


FIG. 12

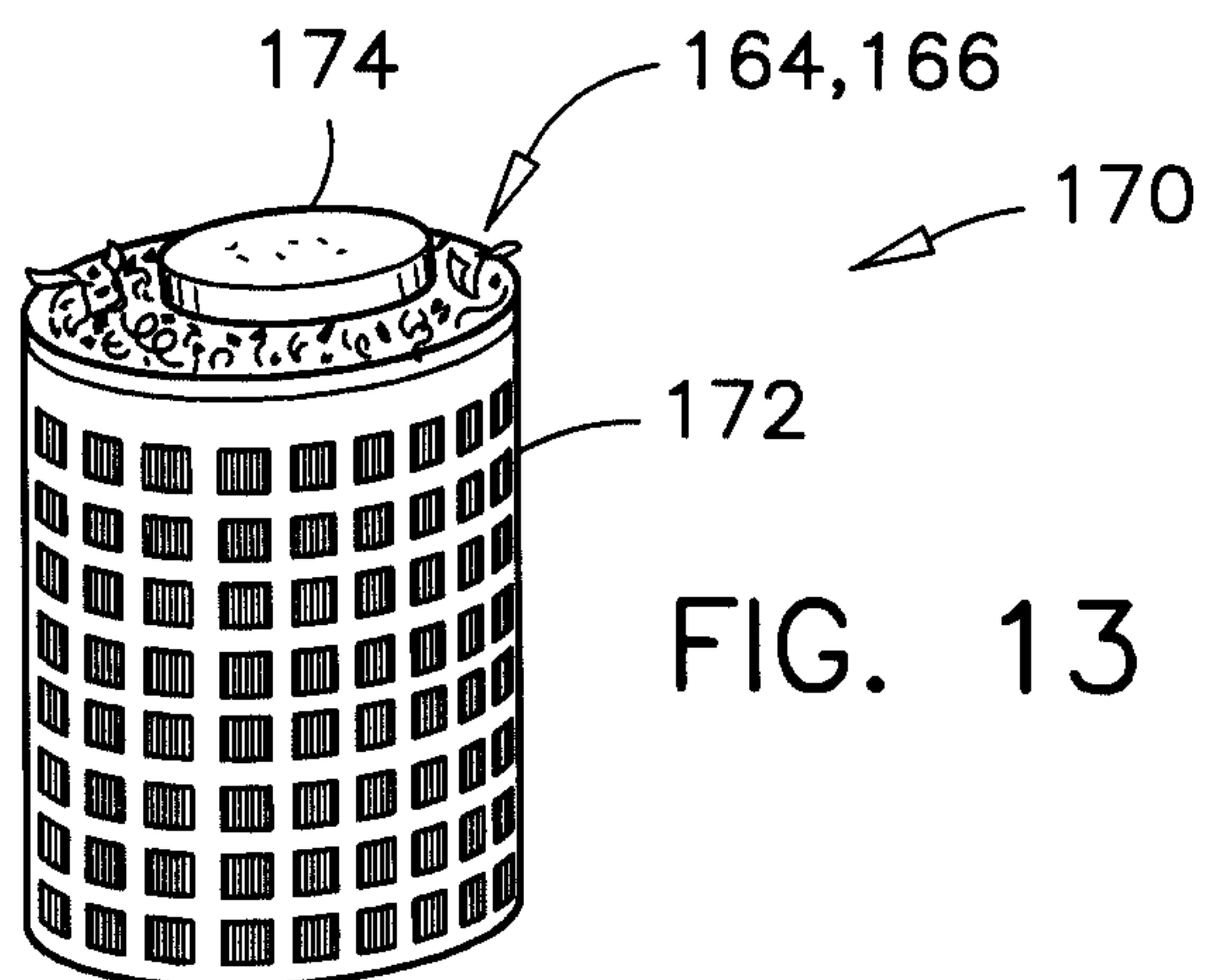
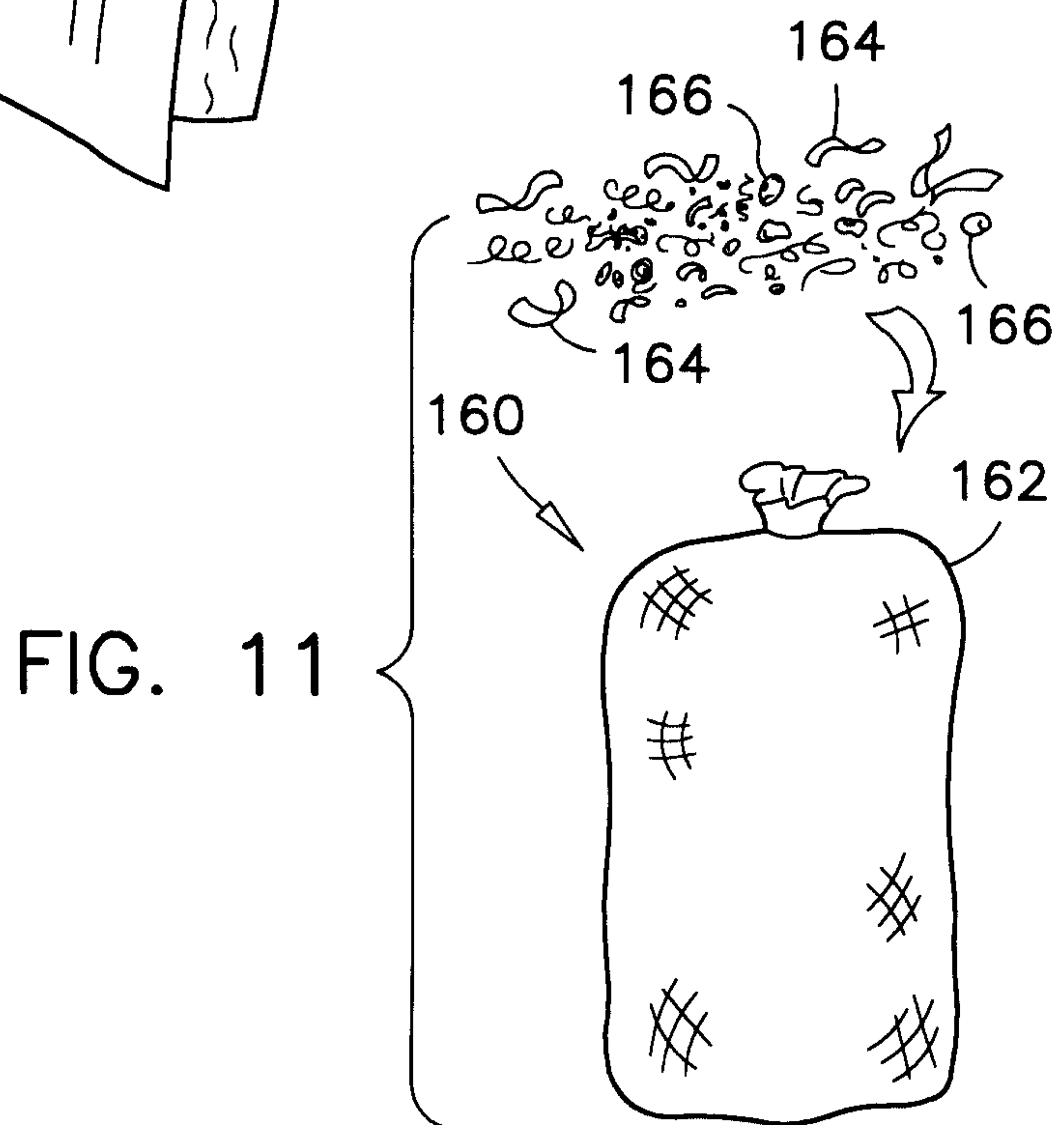


FIG. 13

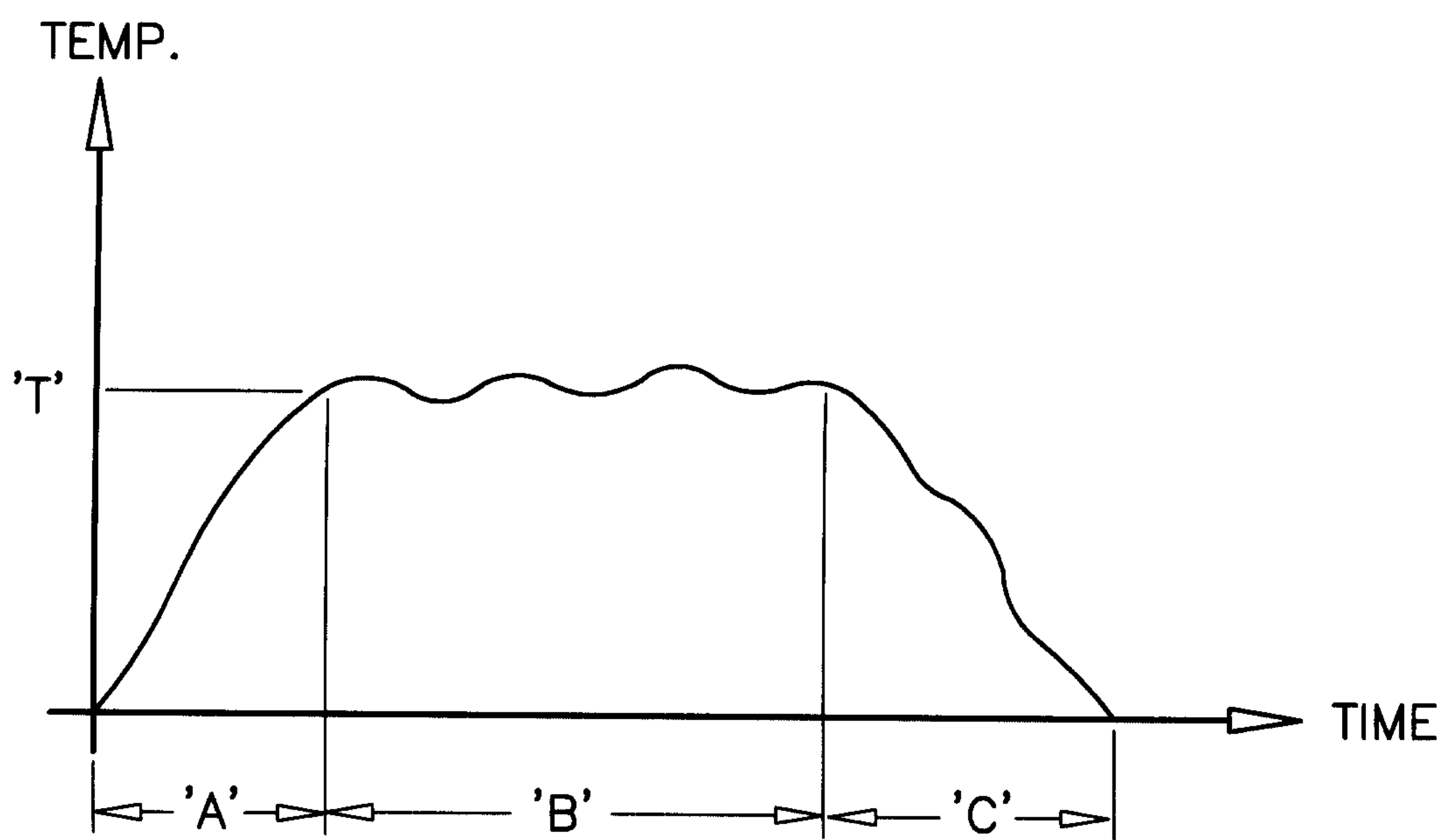


FIG. 14

